



# RESEARCH MEMORANDUM

CORRELATION OF TURBULENT HEAT TRANSFER IN A TUBE FOR

THE DISSOCIATING SYSTEM  $N_2O_4 \rightleftharpoons 2NO_2$

By Richard S. Brokaw

Lewis Flight Propulsion Laboratory  
Cleveland, Ohio

NATIONAL ADVISORY COMMITTEE  
FOR AERONAUTICS

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SUMMARY

Experimental data on turbulent heat transfer in a tube for the dissociating system  $N_2O_4 \rightleftharpoons 2NO_2$  are in accord with the usual Nusselt-Prandtl-Reynolds number correlation for this type of convective heat transfer.

The thermal conductivities and kinematic viscosities needed to formulate the Nusselt, Reynolds, and Prandtl numbers have been computed using rigorous expressions from the kinetic theory of gases. The thermal conductivity includes a large contribution arising from the diffusional transport of chemical enthalpy.

Three methods of selecting average property values for the boundary layer were tested. All gave good correlation using the equilibrium transport properties.

An approximate method using frozen gas properties gave somewhat poor results, but should be useful for complex systems where exact computation would be prohibitively tedious. This approximate method should also be applicable to systems in which chemical equilibrium is not attained, provided the enthalpy of the gases in the free stream and at the wall is known.

INTRODUCTION

Currently there is considerable interest in heat transfer in dissociating gases, because such gases may exist in the exhaust nozzles of jet engines or within the boundary layer on hypersonic aircraft. In reacting gases, heat transport may be considerably higher than in "frozen" (non-reacting) mixtures. Large amounts of heat may be carried as chemical enthalpy of molecules which diffuse because of concentration gradients. These gradients arise because the gas composition varies with temperature.

For example, in a gas which absorbs heat by dissociating as the temperature is raised, heat is transferred when a molecule dissociates in the high-temperature region and diffuses toward the low-temperature region (there is a lower concentration of dissociated molecules at low temperature). In the low-temperature region the gas reassociates, releasing the heat absorbed from the high-temperature dissociation.

Heat transfer in practical systems with chemical reaction has been theoretically examined (e.g., refs. 1 and 2). Parallel theoretical studies of the thermal conductivity of reacting gases have been made: in reference 3 a general expression for the thermal conductivity of gas mixtures in chemical equilibrium is developed, while in reference 4 systems with finite reaction rates are considered.

Recently Beal and Lyerly (ref. 5) have experimentally investigated turbulent heat transfer in a tube for the dissociating system  $N_2O_4 \rightleftharpoons 2NO_2$ . Essential features of their apparatus included a 3/4-inch-diameter test section with a rounded entrance. Heat was added at a thermally guarded hot spot 23 diameters from the tube entry. The guard heater was short, so that the length-diameter ratio for the thermal boundary layer was only 2. Tests were carried out with air in the Reynolds number range from 2500 to 34,000 and with  $N_2O_4 \rightleftharpoons 2NO_2$  mixtures in the range from 18,000 to 104,000.

Beal and Lyerly were unable to analyze their results fully, because of the dearth of experimental conductivity and viscosity data for the  $N_2O_4 \rightleftharpoons 2NO_2$  system in the literature. This report shows that the data of Beal and Lyerly are in complete accord with the usual Nusselt-Prandtl-Reynolds number correlations for this type of convective heat transfer. Thermal conductivities and viscosities for the  $N_2O_4 \rightleftharpoons 2NO_2$  system (needed to formulate the Nusselt, Prandtl, and Reynolds numbers) have been computed using rigorous kinetic theory expressions.

## SYMBOLS

The following symbols are used in this report:

- A dimensionless constant, eq. (1)
- $c_p$  specific heat at constant pressure
- D binary diffusion coefficient
- d tube diameter
- G mass velocity

$\Delta h$  specific enthalpy difference between wall and fluid bulk  
h heat-transfer coefficient  
k thermal conductivity  
Nu Nusselt number,  $h d/k$   
Pr Prandtl number,  $c_p \mu/k$   
Re Reynolds number,  $Gd/\mu$  or  $Gd/\rho v$   
T absolute temperature  
 $\delta$   $\rho D c_{p,f} / k_f$  (dimensionless)  
 $\mu$  absolute viscosity  
 $\nu$  kinematic viscosity  
 $\rho$  density  
 $\varphi$  generalized temperature dependent property

## Subscripts:

B bulk fluid  
e chemical equilibrium  
f chemically frozen  
i contribution due to internal energy  
m value for mixture of monatomic gases  
r contribution due to chemical reaction  
w wall

## Superscripts:

- mean value  
\* evaluated at reference condition

## METHODS OF ANALYSIS

## Transport Properties

The computed thermal conductivities, kinematic viscosities, and Prandtl numbers for the  $N_2O_4 \rightleftharpoons 2NO_2$  system at 1 atmosphere are shown in figure 1. The methods of computation are detailed in the appendix. Calculated conductivities agree with recent experimental determinations (ref. 6), which are also plotted in figure 1. Frozen thermal conductivities and Prandtl numbers are shown by the dashed lines. (The frozen values refer to equilibrium gas composition without chemical reaction.) The horizontal bars across the top of the figure indicate the temperature differences between the wall and bulk fluid for the individual  $N_2O_4 \rightleftharpoons 2NO_2$  heat-transfer tests of Beal and Lyerly; approximate Reynolds numbers are given adjacent to the bars. (The exact values of the Reynolds numbers depend on the way mean property values are chosen.) Over the temperature range of the heat-transfer experiments, the conductivity varies by a factor of 2.2, the kinematic viscosity by a factor of 2.3, and the Prandtl number by a factor of 1.4. These variations are much larger than those encountered with nonreacting gases; over a comparable temperature range, conductivity might increase by 10 to 15 percent and kinematic viscosity by 30 to 40 percent, while Prandtl number would remain substantially constant. Accordingly, for the  $N_2O_4 \rightleftharpoons 2NO_2$  system, suitable mean property values for correlation must be selected.

## Selection of Mean Property Values

In order to formulate the Reynolds, Nusselt, and Prandtl numbers, mean values of viscosity, thermal conductivity, and Prandtl number are required. In actual fact, a mean value of kinematic viscosity is required in the Reynolds number, which may be written  $Gd/\bar{\mu} = Gd/\rho_B \bar{v}$ ; hence,

$$\bar{\mu} = \rho_B \bar{v}$$

Mean properties obtained by the following three methods were tested:

(1) Temperature-integrated properties. If a property is designated as  $\phi$ , then the integrated average value is

$$\bar{\phi} = \frac{\int_{T_B}^{T_W} \phi(T) dT}{T_W - T_B}$$

This type of average suggests itself naturally for thermal conductivity because it is the proper average for heat transfer in the absence of convection (e.g., thermal-conductivity cells).

(2) Reference-temperature properties: Properties are taken at a reference temperature, as suggested in reference 7.

$$T^* = T_B + 0.58 (T_w - T_B)$$

(3) Reference-enthalpy properties: Properties are evaluated at a temperature corresponding to a reference enthalpy  $h^*$ , defined as

$$h^* = h_B + 0.58 (h_w - h_B)$$

### COMPARISON OF CORRELATIONS

#### Correlations Using Equilibrium Transport Properties

All three methods of averaging properties give good correlation of the air and  $N_2O_4 \rightleftharpoons 2NO_2$  heat-transfer data of Beal and Lyerly. Figure 2 shows a log-log plot of  $Nu/Pr^{1/3}$  against Reynolds number for temperature-integrated average properties. A mean kinematic viscosity is required in formulating the Reynolds number. If a mean value of the absolute viscosity is used, correlation is very poor, with errors of up to fifty percent in the Reynolds number. This is shown in figure 3, a reference-enthalpy equilibrium-property correlation, in which a mean absolute viscosity was used. The experimental data for each of the three correlations using mean kinematic viscosity were fitted by an equation of the form

$$Nu/Pr^{1/3} = A Re^{0.79} \quad (1)$$

or

$$\log Nu - \frac{1}{3} \log Pr = \log A + 0.79 \log Re \quad (1a)$$

For each correlation, the constant  $A$  was chosen to give a least-squares fit of all the data (air and  $N_2O_4$ ) at Reynolds numbers greater than 9000 to equation (1a). (The air points at  $Re < 9000$  were excluded because they lie in the transition region.) Results of the correlations are summarized in the first three columns of table I<sup>1</sup>.

<sup>1</sup>It might be noted that in the Reynolds number range under consideration (9000 to 110,000) equation (1), with the  $A$  values of table I, agrees to within 4 percent with Colburn's correlation (ref. 8):  $Nu/Pr^{1/3} = 0.023 Re^{0.8}$ . This is somewhat surprising, in view of the short length-diameter ratio for the thermal boundary layer in Beal and Lyerly's apparatus.

From table I it is seen that all three correlations represent the  $N_2O_4 \rightleftharpoons 2NO_2$  data very well. (The standard deviations for the air data and over-all correlations merely indicate the consistency of the correlations.) For engineering heat-transfer calculations the reference-enthalpy or reference-temperature methods would ordinarily be used, because with these methods, transport properties need be calculated at only one condition.

However, table I shows that the integrated property correlation is slightly better. This is merely a tentative conclusion, however, since the differences between the correlations are largely the result of a single data point (albeit the point with the largest temperature difference). More experimental measurements with larger temperature differences are required to firmly establish the most reliable method for computing mean properties. In any event, for all three correlation methods, deviations for  $N_2O_4$  and air data points are similar. This fact strongly supports the notion that heat-transfer correlations for nonreacting systems may be applied to dissociating systems, provided that transport properties are properly computed and suitable mean values are chosen.

#### Approximate Correlation Using Frozen Transport Properties

In reference 3 it is shown that the equilibrium and frozen thermal conductivities may be related to the equilibrium and frozen specific heats by the equation

$$\frac{k_e}{k_f} = 1 + \delta \left( \frac{c_{p,e}}{c_{p,f}} - 1 \right)$$

If the dimensionless group  $\delta$  is assumed to be unity, then  $k_e \approx k_f c_{p,e} / c_{p,f}$ . (Aerodynamicists sometimes refer to  $\delta$  as the "Lewis number." However, this usage is not generally accepted.) Hence, for a system with a finite temperature difference, it seems reasonable to define an approximate mean thermal conductivity by

$$\bar{k}_e \approx \left( \frac{\bar{k}_f}{c_{p,f}} \right) \frac{\int_{T_B}^{T_w} c_{p,e} dT}{T_w - T_B} = \bar{k}_f \frac{\Delta h}{c_{p,f} \Delta T} \quad (2)$$

with the frozen specific heat and conductivity evaluated at a reference temperature. Subject to this approximation, the Nusselt and Prandtl numbers are

$$Nu_e \approx Nu_f \frac{\bar{c}_{p,f} \Delta T}{\Delta h}$$

$$Pr_e \approx Pr_f$$

A reference-temperature correlation using frozen properties is shown in figure 4. (If the correction term  $\bar{c}_{p,f}\Delta T/\Delta h$  were not applied to the frozen Nusselt number, Nusselt numbers for the  $N_2O_4 \rightleftharpoons 2NO_2$  system would be eight- to ninefold too high.) As expected, the correlation of figure 4 is poorer than the correlation of figure 2; this is reflected in a higher standard deviation, shown in the fourth column of table I. In reality the deviations are systematic with  $\delta$  and might amount to 10 percent or more under some conditions. For this system  $\delta$  actually ranges from 1.13 at  $310^\circ K$  to 0.64 at  $375^\circ K$ . The frozen property correlation can be improved by introducing a correction term involving  $\delta$  and plotting

$$\left( \frac{Nu_f}{Pr_f^{1/3}} \right) \left( \bar{c}_{p,f}\Delta T/\Delta h \right) \left[ 1 + (\delta - 1) \left( 1 - \frac{\bar{c}_{p,f}\Delta T}{\Delta h} \right) \right]^{-2/3}$$

against Reynolds number. The correction term is reminiscent of similar corrections used in treatments of hypersonic stagnation-point heat transfer (refs. 2 and 9). However, heat transfer in multicomponent mixtures cannot, in general, be characterized by a single diffusion coefficient, and thus the correction term is not generally useful<sup>2</sup>.

The approximate correlation without the correction should be applicable to many complex systems where an exact computation would be prohibitively tedious (e.g., jet-engine and rocket exhausts). Indeed, it should be applicable to systems in which chemical equilibrium is not attained, provided the chemical compositions (and hence, enthalpy) of the gases in the free stream and at the wall are known.

### CONCLUSIONS

Experimental data for turbulent heat transfer in a tube have been analyzed and correlated. The following conclusions may be drawn:

1. Experimental heat-transfer data in a tube for the dissociating system  $N_2O_4 \rightleftharpoons 2NO_2$  correlate well with air data using transport properties computed on the basis of chemical equilibrium.
2. In computing Reynolds number a mean kinematic viscosity is required.

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<sup>2</sup>For multicomponent mixtures  $\delta$  may be defined as  $[(k_e/k_f) - 1]/[(c_{p,e}/c_{p,f}) - 1]$ , and the correction term is valid. This is not useful, however, since the equilibrium properties calculated to obtain  $\delta$  might better be applied directly in the correlations of the preceding section.

3. The best correlation was obtained using temperature-integrated properties. Reference-temperature and reference-enthalpy properties were almost as good.

4. Errors are somewhat larger for an approximate correlation which uses frozen properties. In this method the frozen Nusselt number is corrected by the factor  $c_{p,f}\Delta T/\Delta h$  and it is assumed that  $\rho D c_{p,f}/k_f = 1$ . (Mean value of the specific heat at constant pressure and chemically frozen conditions,  $\bar{c}_{p,f}$ ; temperature difference between wall and fluid bulk,  $\Delta T$ ; specific enthalpy difference between wall and fluid bulk,  $\Delta h$ ; density,  $\rho$ ; binary diffusion coefficient,  $D$ ; specific heat at constant pressure and chemically frozen conditions,  $c_{p,f}$ ; chemically frozen thermal conductivity,  $k_f$ .)

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## APPENDIX - CALCULATION OF EQUILIBRIUM PROPERTIES

FOR THE  $N_2O_4 \rightleftharpoons 2NO_2$  SYSTEM

Transport-property calculations were carried out assuming a Lennard-Jones (6-12) potential for the molecular interactions (ref. 10). Gas compositions were obtained from the equilibrium data of Bodenstein (ref. 11). The means by which various properties were computed are as follows:

Force constants. - In order to compute the transport properties, force constants  $\sigma$  and  $\epsilon/k$  characteristic of molecular interactions must be estimated. For the  $N_2O_4$  molecule, the force constants were estimated from critical temperature and density (ref. 10, eqs. (4.1-17) and (4.1-18)). The critical temperature and density for pure  $N_2O_4$  were in turn estimated from the boiling point and density of liquid  $N_2O_4$  by the methods of references 12 and 13. (It was assumed that the boiling point and liquid density of  $N_2O_4 \rightleftharpoons 2NO_2$  equilibrium mixtures is close to the value which would be expected for pure  $N_2O_4$ .) Force constants chosen for  $N_2O_4$  were:  $\sigma = 4.74 \text{ \AA}$  and  $\epsilon/k = 383^\circ \text{ K}$ .

The force constants for the  $NO_2$  molecule were assumed to be similar to those of  $CO_2$  and  $N_2O$  (ref. 10, p. 1111); values used were  $\sigma = 3.9 \text{ \AA}$  and  $\epsilon/k = 230^\circ \text{ K}$ .

For the  $N_2O_4$ - $NO_2$  interaction,  $\sigma$  was taken as the arithmetic mean of  $NO_2$  and  $N_2O_4$  values;  $\epsilon/k$  was taken as the geometric mean (ref. 10, eqs. (8.4-8) and (8.4-9)).

These force constants differ somewhat from those used in reference 3; they are believed to be more realistic estimates.

Viscosities. - Viscosities were calculated by equation (8.2-22) of reference 10.

Thermal conductivities. - The thermal conductivity of a reacting gas mixture may be broken into three portions:

$$k_e = k_m + k_i + k_r \quad (3)$$

The first term on the right side of equation (3) represents collisional kinetic energy transport; it was computed using equation (8.2-36) of reference 10. The second term (in the nature of an Eucken correction) represents a diffusional transport of internal energy (rotation and vibration) and was calculated using equation (75) of reference 14. The third

and dominant term for these dissociating mixtures is the increase in conductivity due to chemical reaction. It was computed from equation (14) of reference 3. The frozen thermal conductivity is simply the first two terms of equation (3).

Diffusion coefficients. - Diffusion coefficients were computed according to equation (8.2-44) of reference 10.

Specific heats. - The specific heat may be broken up into a frozen part and a contribution due to chemical reaction. The frozen specific heats were computed from the known vibrational frequencies of the  $\text{NO}_2$  and  $\text{N}_2\text{O}_4$  molecules. The contribution due to chemical reaction was computed according to reference 3.

Heat of reaction. - The heat of reaction at  $298.16^\circ\text{K}$  was computed from the heats of formation of the  $\text{NO}_2$  and  $\text{N}_2\text{O}_4$  molecules presented in reference 15. Values at other temperatures were obtained by correcting the  $298.16^\circ\text{K}$  value for the integrated heat capacity difference between products and reactant.

Calculated values of viscosity, density divided by pressure, frozen and equilibrium thermal conductivity, and frozen and equilibrium specific heat are contained in table II. Values were computed at  $10^\circ$  intervals from  $290^\circ$  to  $490^\circ\text{K}$  and for pressures of 1.0, 0.74, and 0.33 atmospheres. (Additional thermal conductivities computed for pressures of 0.5, 0.2, 0.1, 0.05, and 0.02 atmospheres may be found in reference 6.) Only the 1.0-atmosphere values were used in the correlations of this report.

#### REFERENCES

1. Altman, D., and Wise, H.: Effect of Chemical Reactions in the Boundary Layer on Convective Heat Transfer. *Jet Prop.*, vol. 26, no. 4, Apr. 1956, pp. 256-258; 269.
2. Fay, J. A., Riddell, F. R., and Kemp, N. H.: Stagnation Point Heat Transfer in Dissociated Air Flow. *Jet Prop.*, vol. 27, no. 6, June 1957, pp. 672-674.
3. Butler, J. N., and Brokaw, R. S.: Thermal Conductivity of Gas Mixtures in Chemical Equilibrium. *Jour. Chem. Phys.*, vol. 26, no. 6, June 1957, pp. 1636-1643.
4. Hirschfelder, Joseph O.: Heat Transfer in Chemically Reacting Mixtures, I. *Jour. Chem. Phys.*, vol. 26, no. 2, Feb. 1957, pp. 274-281.

5. Beal, John L., and Lyerly, Ray L.: The Influence of Gas Dissociation on Heat Transfer. WADC Tech. Rep. 56-494, Wright Air Dev. Center, Wright-Patterson Air Force Base, Sept. 1956. (Contract AF 33(616)-2954.) (ASTIA No. AD 118074.)
6. Coffin, Kenneth P., and O'Neal, Cleveland, Jr.: Experimental Thermal Conductivities of the  $N_2O_4 \rightleftharpoons 2NO_2$  System. NACA TN 4209, 1957.
7. Johnson, H. A., and Rubesin, M. W.: Aerodynamic Heating and Convective Heat Transfer - Summary of Literature Survey. Trans. ASME, vol. 71, no. 5, July 1949, pp. 447-456.
8. Colburn, Allan P.: A Method of Correlating Forced Convection Heat Transfer Data and a Comparison with Fluid Friction. Trans. Am. Inst. Chem. Eng., vol. XXIX, 1933, pp. 174-210.
9. Lees, L.: Laminar Heat Transfer over Blunt-Nosed Bodies at Hypersonic Flight Speeds. Jet Prop., vol. 26, no. 4, Apr. 1956, pp. 259-269; 274.
10. Hirschfelder, Joseph O., Curtiss, Charles F., and Bird, R. Byron: Molecular Theory of Gases and Liquids. John Wiley & Sons, Inc., 1954.
11. Bodenstein, Max: Bildung und Zersetzung der höheren Stickoxyde. Zs. f. Phys. Chem., Bd. 100, 1922, pp. 68-123.
12. Hougen, Olaf A., and Watson, Kenneth M.: Chemical Process Principles. Pt. 1. Material and Energy Balances. John Wiley & Sons, Inc., 1943, p. 69.
13. Watson, K. M.: Thermodynamics of the Liquid State - Generalized Prediction of Properties. Ind. and Eng. Chem., vol. 35, no. 4, Apr. 1943, pp. 398-406.
14. Hirschfelder, Joseph O.: Heat Conductivity in Polyatomic, Electronically Excited, or Chemically Reacting Mixtures, III. Sixth Symposium (International) on Combustion, Reinhold Pub. Corp., 1957, pp. 351-366.
15. Rossini, Frederick D., et al.: Selected Values of Chemical Thermo-dynamic Properties. Cir. 500, NBS, Feb. 1952.

TABLE I. - COMPARISON OF HEAT-TRANSFER CORRELATIONS

	Integrated properties	Reference-temperature properties	Reference-enthalpy properties	Reference-temperature frozen properties, $\delta = 1$
A (eq. 1(a))	0.0261	0.0262	0.0257	0.0262
Standard deviation				
$\text{N}_2\text{O}_4 \neq 2\text{NO}_2$ data	2.4 percent	3.3 percent	3.5 percent	5.3 percent
Air ( $Re > 9000$ )	3.0 percent	3.1 percent	3.2 percent	-----
Over-all ( $Re > 9000$ )	2.9 percent	3.1 percent	3.3 percent	-----

TABLE II. - COMPUTED PROPERTIES FOR  $N_2O_4 \rightleftharpoons 2NO_2$  EQUILIBRIUM  
(a) Pressure, 1 atmosphere.

Temperature, °K	Viscosity, poises	Equilibrium thermal conductivity, cal/(cm)(sec)(°K)	Frozen thermal conductivity,(°K)	Density + pressure, gm/(cm <sup>3</sup> )(atm)	Equilibrium specific heat, cal/gm	Frozen specific heat, cal/gm
290	$1.124 \times 10^{-4}$	$20.083 \times 10^{-5}$	$3.110 \times 10^{-5}$	$3.414 \times 10^{-3}$	1.0168	0.2017
300	1.192	25.668	3.337	3.130	1.2973	.2042
310	1.270	31.176	3.587	2.832	1.6051	.2060
320	1.356	35.233	3.860	2.540	1.8730	.2075
330	1.445	36.366	4.143	2.272	2.0067	.2083
340	1.534	33.977	4.427	2.043	1.9395	.2089
350	1.615	29.086	4.693	1.864	1.6893	.2094
360	1.689	23.308	4.938	1.726	1.3483	.2099
370	1.754	18.037	5.163	1.621	1.0158	.2106
380	1.811	14.087	5.366	1.542	.7556	.2115
390	1.865	11.262	5.561	1.479	.5650	.2125
400	1.915	9.478	5.745	1.428	.4415	.2138
410	1.961	8.375	5.923	1.385	.3622	.2151
420	2.005	7.742	6.098	1.346	.3134	.2166
430	2.048	7.397	6.272	1.312	.2831	.2181
440	2.091	7.219	6.446	1.280	.2634	.2196
450	2.133	7.143	6.617	1.250	.2503	.2211
460	2.174	7.163	6.791	1.221	.2429	.2226
470	2.215	7.223	6.958	1.195	.2382	.2240
480	2.254	7.315	7.127	1.169	.2353	.2255
490	2.294	7.436	7.300	1.145	.2341	.2270

(b) Pressure, 0.74 atmosphere.

Temperature, °K	Viscosity, poises	Equilibrium thermal conductivity, cal/(cm)(sec)(°K)	Frozen thermal conductivity,(°K)	Density + pressure, gm/(cm <sup>3</sup> )(atm)	Equilibrium specific heat, cal/gm	Frozen specific heat, cal/gm
290	$1.134 \times 10^{-4}$	$22.332 \times 10^{-5}$	$3.136 \times 10^{-5}$	$3.352 \times 10^{-3}$	1.1407	0.2015
300	1.206	28.201	3.373	3.053	1.4500	.2039
310	1.288	30.498	3.633	2.745	1.7685	.2056
320	1.377	36.578	3.916	2.449	2.0051	.2069
330	1.468	36.073	4.203	2.189	2.0543	.2076
340	1.554	32.025	4.482	1.974	1.8776	.2081
350	1.633	26.197	4.740	1.811	1.5482	.2086
360	1.702	20.304	4.975	1.688	1.1803	.2092
370	1.763	15.542	5.188	1.596	.8677	.2101
380	1.817	12.201	5.384	1.526	.6410	.2111
390	1.870	9.939	5.573	1.469	.4841	.2123
400	1.918	8.576	5.753	1.422	.3866	.2136
410	1.963	7.769	5.929	1.380	.3258	.2150
420	2.006	7.331	6.102	1.343	.2890	.2164
430	2.049	7.114	6.275	1.310	.2666	.2180
440	2.092	7.024	6.448	1.278	.2522	.2196
450	2.133	7.008	6.618	1.249	.2427	.2210
460	2.175	7.069	6.792	1.221	.2377	.2226
470	2.215	7.154	6.959	1.194	.2344	.2240
480	2.254	7.267	7.128	1.169	.2328	.2255
490	2.294	7.401	7.300	1.145	.2322	.2270

(c) Pressure, 0.33 atmosphere.

Temperature, °K	Viscosity, poises	Equilibrium thermal conductivity, cal/(cm)(sec)(°K)	Frozen thermal conductivity,(°K)	Density + pressure, gm/(cm <sup>3</sup> )(atm)	Equilibrium specific heat, cal/gm	Frozen specific heat, cal/gm
290	$1.166 \times 10^{-4}$	$29.106 \times 10^{-5}$	$3.221 \times 10^{-5}$	$3.152 \times 10^{-3}$	1.5476	0.2010
300	1.250	34.829	3.488	2.819	1.9064	.2030
310	1.342	37.844	3.776	2.497	2.1643	.2043
320	1.435	36.481	4.072	2.216	2.1774	.2051
330	1.523	31.189	4.352	1.994	1.9129	.2057
340	1.600	24.217	4.608	1.828	1.4900	.2063
350	1.667	18.109	4.835	1.709	1.0850	.2070
360	1.727	13.554	5.042	1.621	.7694	.2080
370	1.780	10.557	5.234	1.554	.5566	.2092
380	1.828	8.752	5.414	1.499	.4248	.2105
390	1.877	7.663	5.593	1.452	.3422	.2118
400	1.922	7.078	5.766	1.411	.2943	.2133
410	1.966	6.781	5.937	1.373	.2658	.2148
420	2.009	6.667	6.108	1.339	.2494	.2163
430	2.051	6.657	6.279	1.307	.2398	.2179
440	2.093	6.711	6.451	1.276	.2342	.2195
450	2.134	6.796	6.620	1.247	.2307	.2210
460	2.175	6.918	6.794	1.220	.2293	.2225
470	2.215	7.048	6.960	1.194	.2287	.2239
480	2.255	7.191	7.129	1.169	.2287	.2254
490	2.294	7.346	7.301	1.145	.2293	.2270

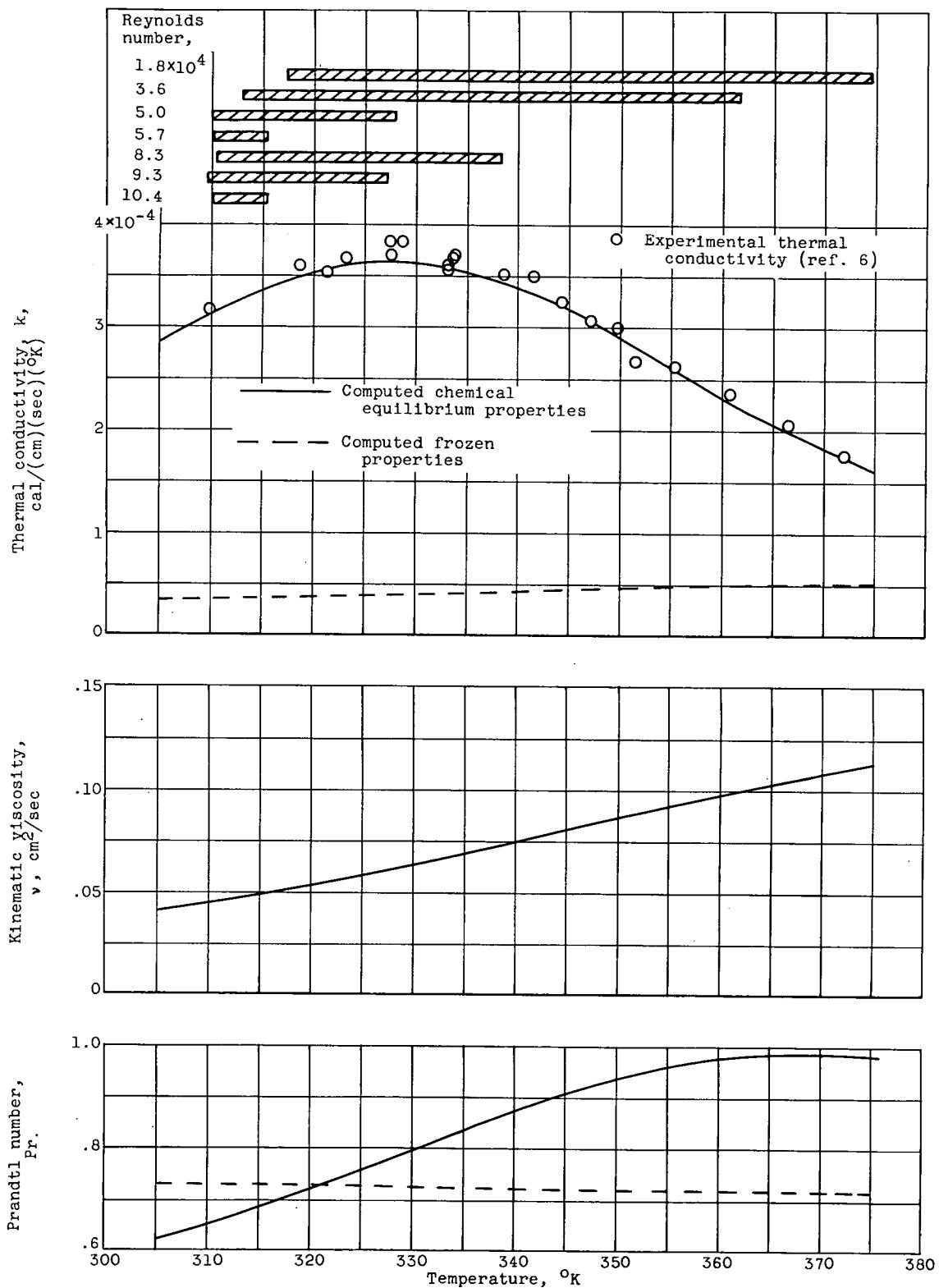


Figure 1. - Equilibrium transport properties for  $\text{N}_2\text{O}_4 + 2\text{NO}_2$  system at one atmosphere.

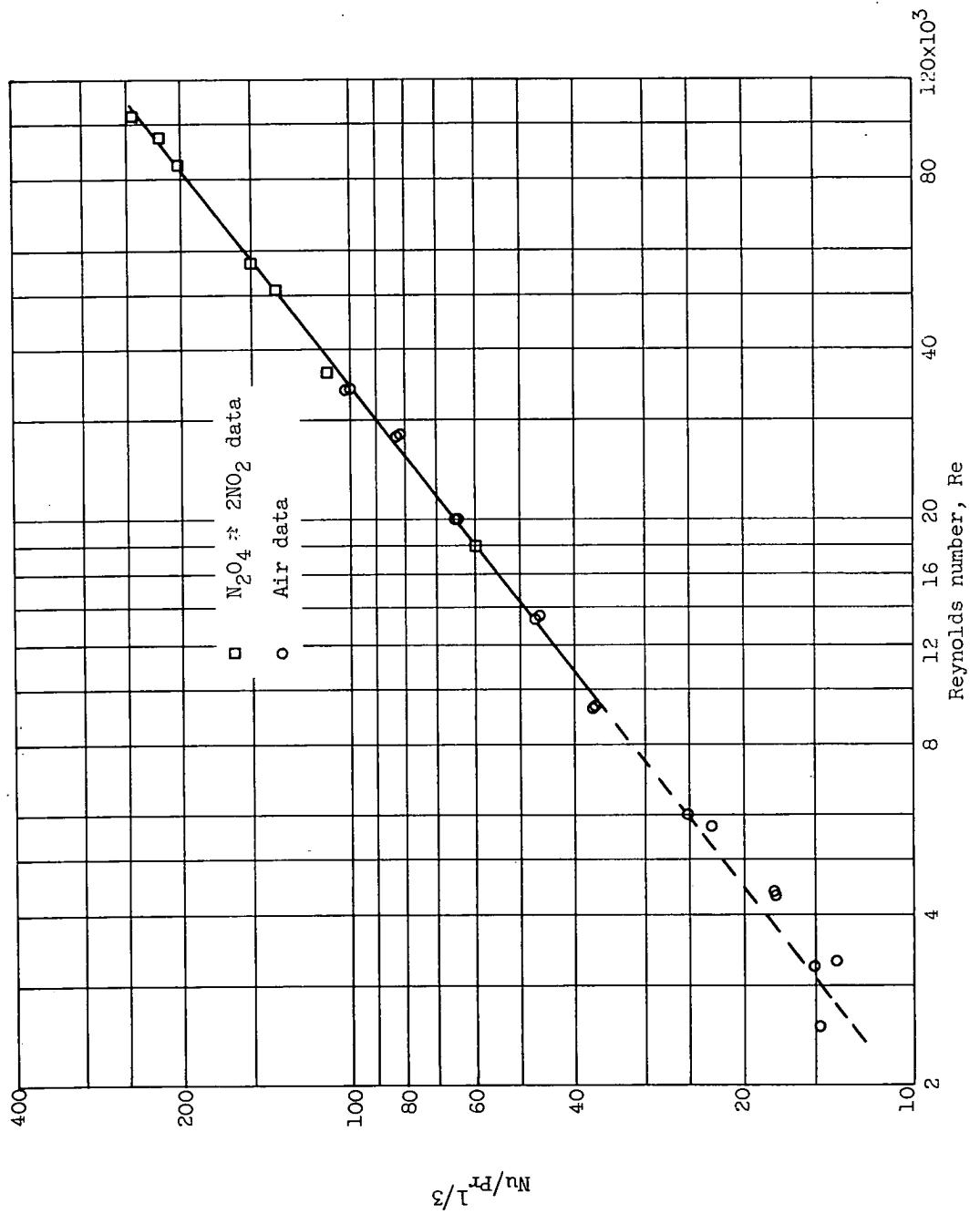


Figure 2. - Correlation of data with temperature-integrated equilibrium properties.  
(Mean kinematic viscosity used to formulate Reynolds number.)

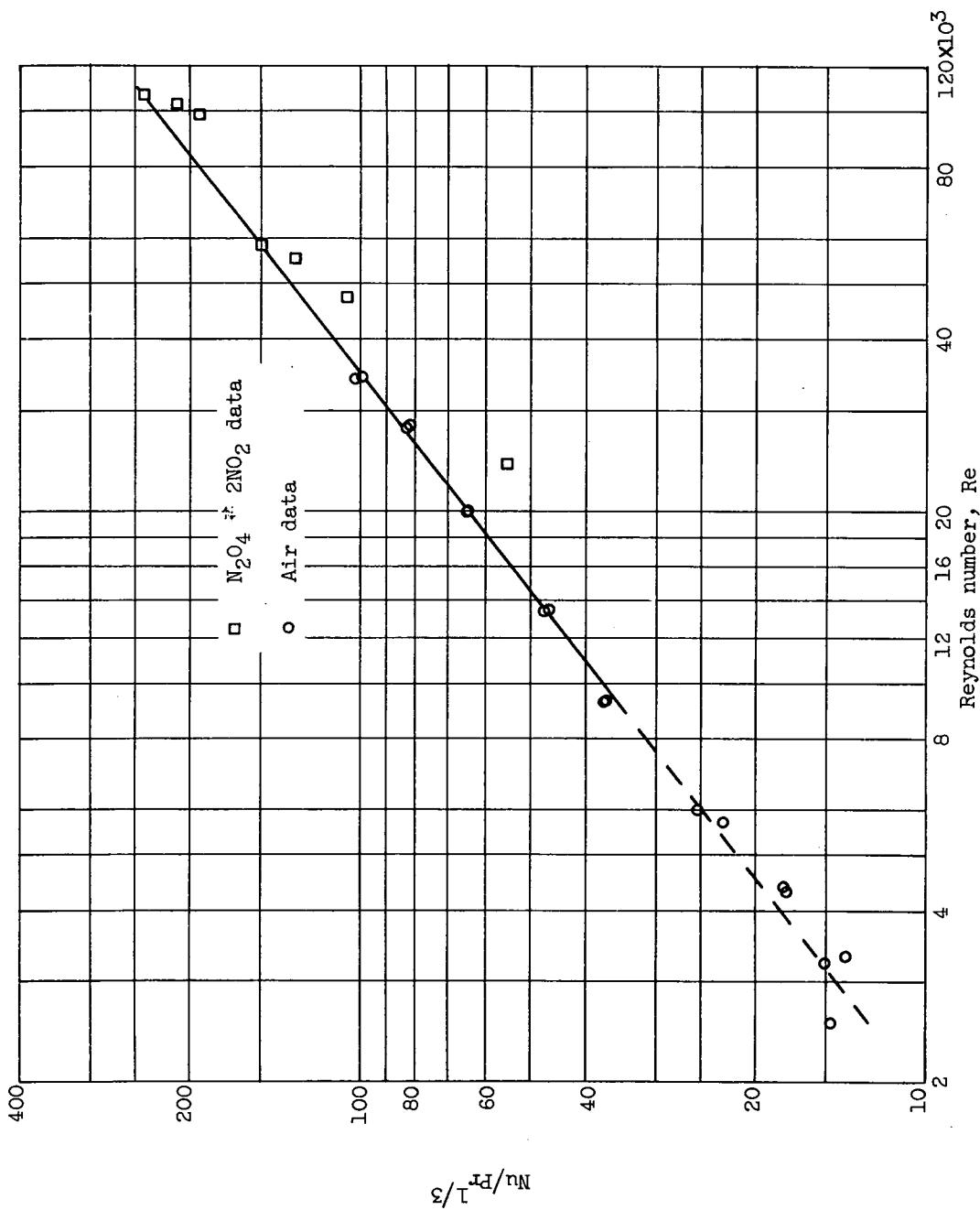


Figure 3. - Correlation of data with reference-enthalpy equilibrium properties.  
(Mean absolute viscosity used to formulate Reynolds number.)

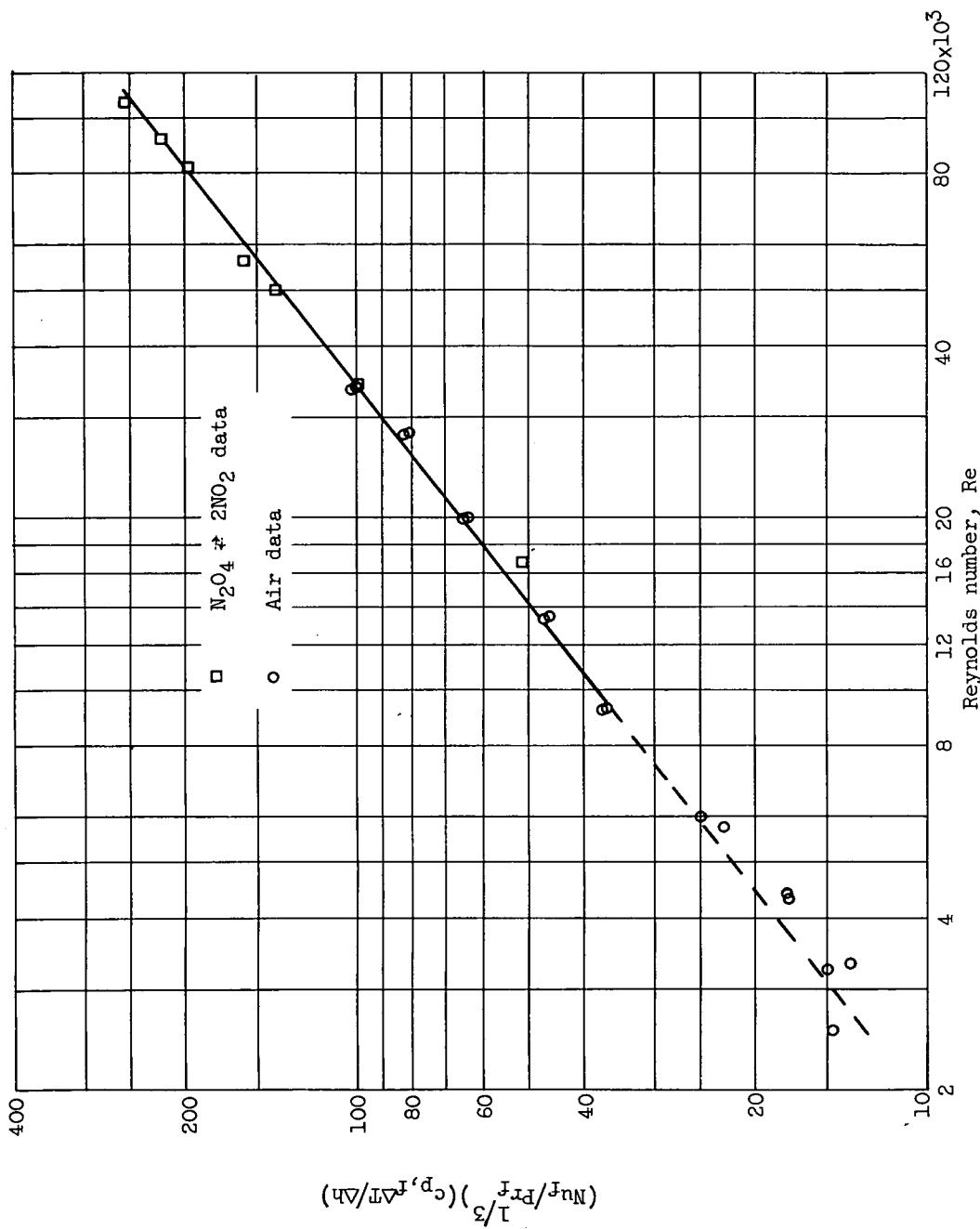


Figure 4. - Approximate correlation of data with reference-temperature frozen properties.