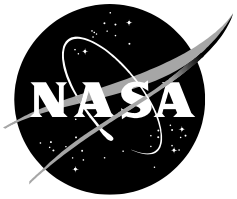
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Curve Fits of the NIST-JANAF Thermochemical Tables Fourth Edition

Timothy K. Risch
Armstrong Flight Research Center, Edwards, California

February 2021

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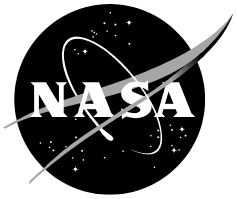
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Ba_1Br_2(cr) Barium Bromide (BaBr2)....F-261
Ba_1Br_2(cr,l) Barium Bromide (BaBr2)....F-262
Ba_1Br_2(g) Barium Bromide (BaBr2)....F-263
Ba_1Br_2(l) Barium Bromide (BaBr2)....F-264
Ba_1Cl_1(g) Barium Chloride (BaCl)....F-265
Ba_1Cl_2(cr) Barium Chloride (BaCl2)....F-266
Ba_1Cl_2(cr,l) Barium Chloride (BaCl2)....F-267
Ba_1Cl_2(g) Barium Chloride (BaCl2)....F-268
Ba_1Cl_2(l) Barium Chloride (BaCl2)....F-269
Ba_1F_1(g) Barium Fluoride (BaF)....F-270
$Ba_1F_1^+$(g) Barium Fluoride, Ion (BaF+)....F-271
Ba_1F_2(cr) Barium Fluoride (BaF2)....F-272
Ba_1F_2(cr,l) Barium Fluoride (BaF2)....F-273
Ba_1F_2(g) Barium Fluoride (BaF2)....F-274
Ba_1F_2(l) Barium Fluoride (BaF2)....F-275
$Ba_1H_1O_1$(g) Barium Hydroxide (BaOH)....F-276
$Ba_1H_1O_1^+$(g) Barium Hydroxide, Ion (BaOH+)....F-277
$Ba_1H_2O_2$(cr) Barium Hydroxide, Alpha (Ba(OH)2)....F-278
$Ba_1H_2O_2$(cr,l) Barium Hydroxide (Ba(OH)2)....F-279
$Ba_1H_2O_2$(g) Barium Hydroxide (Ba(OH)2)....F-280
$Ba_1H_2O_2$(l) Barium Hydroxide (Ba(OH)2)....F-281
Ba_1I_1(g) Barium Iodide (BaI)....F-282
Ba_1I_2(cr) Barium Iodide (BaI2)....F-283
Ba_1I_2(cr,l) Barium Iodide (BaI2)....F-284
Ba_1I_2(g) Barium Iodide (BaI2)....F-285
Ba_1I_2(l) Barium Iodide (BaI2)....F-286
Ba_1O_1(cr) Barium Oxide (BaO)....F-287
Ba_1O_1(cr,l) Barium Oxide (BaO)....F-288
Ba_1O_1(g) Barium Oxide (BaO)....F-289

Ba_1O_1(l)	Barium Oxide (BaO)	F-290
Ba_1S_1(cr)	Barium Sulfide (BaS)	F-291
Ba_1S_1(g)	Barium Sulfide (BaS)	F-292
Be1(cr)	Beryllium (Be)	F-293
Be1(cr,l)	Beryllium (Be)	F-294
Be1(g)	Beryllium (Be)	F-295
Be1(l)	Beryllium (Be)	F-296
Be1+(g)	Beryllium, Ion (Be^+)	F-297
Be1Br1(g)	Beryllium Bromide (BeBr)	F-298
Be1Br2(cr)	Beryllium Bromide ($BeBr_2$)	F-299
Be1Br2(g)	Beryllium Bromide ($BeBr_2$)	F-300
Be1Cl1(g)	Beryllium Chloride (BeCl)	F-301
Be1Cl1+(g)	Beryllium Chloride, Ion ($BeCl^+$)	F-302
Be1Cl1F1(g)	Beryllium Chloride Fluoride (BeClF)	F-303
Be1Cl2(cr)_01	Beryllium Chloride, Alpha ($BeCl_2$)	F-304
Be1Cl2(cr)_02	Beryllium Chloride, Beta ($BeCl_2$)	F-305
Be1Cl2(cr,l)	Beryllium Chloride ($BeCl_2$)	F-306
Be1Cl2(g)	Beryllium Chloride ($BeCl_2$)	F-307
Be1Cl2(l)	Beryllium Chloride ($BeCl_2$)	F-308
Be1F1(g)	Beryllium Fluoride (BeF)	F-309
Be1F2(cr)	Beryllium Fluoride (BeF_2)	F-310
Be1F2(cr,l)	Beryllium Fluoride (BeF_2)	F-311
Be1F2(g)	Beryllium Fluoride (BeF_2)	F-312
Be1F2(l)	Beryllium Fluoride (BeF_2)	F-313
Be1F3Li1(cr)	Lithium Trifluoroberyllate ($LiBeF_3$)	F-314
Be1F3Li1(cr,l)	Lithium Trifluoroberyllate ($LiBeF_3$)	F-315
Be1F3Li1(g)	Lithium Trifluoroberyllate ($LiBeF_3$)	F-316
Be1F3Li1(l)	Lithium Trifluoroberyllate ($LiBeF_3$)	F-317
Be1F4Li2(cr)	Lithium Tetrafluoroberyllate (Li_2BeF_4)	F-318
Be1F4Li2(cr,l)	Lithium Tetrafluoroberyllate (Li_2BeF_4)	F-319
Be1F4Li2(l)	Lithium Tetrafluoroberyllate (Li_2BeF_4)	F-320
Be1H1(g)	Beryllium Hydride (BeH)	F-321
Be1H1+(g)	Beryllium Hydride, Ion (BeH^+)	F-322
Be1H1O1(g)	Beryllium Hydroxide (BeOH)	F-323
Be1H1O1+(g)	Beryllium Hydroxide, Ion ($BeOH^+$)	F-324
Be1H2(g)	Beryllium Hydride (BeH_2)	F-325
Be1H2O2(cr)_01	Beryllium Hydroxide, Alpha ($Be(OH)_2$)	F-326
Be1H2O2(cr)_02	Beryllium Hydroxide, Beta ($Be(OH)_2$)	F-327
Be1H2O2(g)	Beryllium Hydroxide ($Be(OH)_2$)	F-328
Be1I1(g)	Beryllium Iodide (BeI)	F-329
Be1I2(cr)	Beryllium Iodide (BeI_2)	F-330
Be1I2(cr,l)	Beryllium Iodide (BeI_2)	F-331
Be1I2(g)	Beryllium Iodide (BeI_2)	F-332
Be1I2(l)	Beryllium Iodide (BeI_2)	F-333
Be1N1(g)	Beryllium Nitride (BeN)	F-334
Be1O1(cr)_01	Beryllium Oxide, Alpha (BeO)	F-335

Be1O1(cr)_02	Beryllium Oxide, Beta (BeO)	F-336
Be1O1(cr,l)	Beryllium Oxide (BeO)	F-337
Be1O1(g)	Beryllium Oxide (BeO)	F-338
Be1O1(l)	Beryllium Oxide (BeO)	F-339
Be1O4S1(cr)_01	Beryllium Sulfate, Alpha ($BeSO_4$)	F-340
Be1O4S1(cr)_02	Beryllium Sulfate, Beta ($BeSO_4$)	F-341
Be1O4S1(cr)_03	Beryllium Sulfate, Gamma ($BeSO_4$)	F-342
Be1O4W1(cr)	Beryllium Tungsten Oxide ($BeWO_4$)	F-343
Be1S1(cr)	Beryllium Sulfide (BeS)	F-344
Be1S1(g)	Beryllium Sulfide (BeS)	F-345
Be2(g)	Beryllium (Be_2)	F-346
Be2Cl4(g)	Beryllium Chloride ($(BeCl_2)_2$)	F-347
Be2F2O1(g)	Beryllium Fluoride Oxide ($O(BeF)_2$)	F-348
Be2O1(g)	Beryllium Oxide (Be_2O)	F-349
Be2O2(g)	Beryllium Oxide ($(BeO)_2$)	F-350
Be2O4Si1(cr)	Beryllium Silicate (Be_2SiO_4)	F-351
Be3N2(cr)	Beryllium Nitride, Alpha (Be_3N_2)	F-352
Be3N2(cr,l)	Beryllium Nitride (Be_3N_2)	F-353
Be3N2(l)	Beryllium Nitride (Be_3N_2)	F-354
Be3O3(g)	Beryllium Oxide ($(BeO)_3$)	F-355
Be4O4(g)	Beryllium Oxide ($(BeO)_4$)	F-356
Be5O5(g)	Beryllium Oxide ($(BeO)_5$)	F-357
Be6O6(g)	Beryllium Oxide ($(BeO)_6$)	F-358
Br1(g)	Bromine (Br)	F-359
Br1-(g)	Bromine, Ion (Br^-)	F-360
Br1+(g)	Bromine, Ion (Br^+)	F-361
Br1Ca1(g)	Calcium Bromide (CaBr)	F-362
Br1Cl1(g)	Bromine Chloride (BrCl)	F-363
Br1F1(g)	Bromine Fluoride (BrF)	F-364
Br1F3(g)	Bromine Fluoride (BrF_3)	F-365
Br1F5(g)	Bromine Fluoride (BrF_5)	F-366
Br1F5S1(g)	Sulfur Bromide Fluoride ($SBrF_5$)	F-367
Br1H1(g)	Hydrogen Bromide (HBr)	F-368
Br1H3Si1(g)	Bromosilane (SiH_3Br)	F-369
Br1H4N1(cr)	Ammonium Bromide (NH_4Br)	F-370
Br1Hg1(g)	Mercury Bromide (HgBr)	F-371
Br1I1(g)	Iodine Bromide (IBr)	F-372
Br1K1(cr)	Potassium Bromide (KBr)	F-373
Br1K1(cr,l)	Potassium Bromide (KBr)	F-374
Br1K1(g)	Potassium Bromide (KBr)	F-375
Br1K1(l)	Potassium Bromide (KBr)	F-376
Br1Li1(cr)	Lithium Bromide (LiBr)	F-377
Br1Li1(cr,l)	Lithium Bromide (LiBr)	F-378
Br1Li1(g)	Lithium Bromide (LiBr)	F-379
Br1Li1(l)	Lithium Bromide (LiBr)	F-380
Br1Mg1(g)	Magnesium Bromide (MgBr)	F-381

Br1Mo1(g) Molybdenum Bromide (MoBr) F-382
Br1N1(g) Bromoimidogen (NBr) F-383
Br1N1O1(g) Nitrosyl Bromide (ONBr) F-384
Br1Na1(cr) Sodium Bromide (NaBr) F-385
Br1Na1(cr,l) Sodium Bromide (NaBr) F-386
Br1Na1(g) Sodium Bromide (NaBr) F-387
Br1Na1(l) Sodium Bromide (NaBr) F-388
Br1P1(g) Phosphorus Bromide (PBr) F-389
Br1Pb1(g) Lead Bromide (PbBr) F-390
Br1Si1(g) Bromosilylidyne (SiBr) F-391
Br1Sr1(g) Strontium Bromide (SrBr) F-392
Br1Ti1(g) Titanium Bromide (TiBr) F-393
Br1W1(g) Tungsten Bromide (WBr) F-394
Br1Zr1(g) Zirconium Bromide (ZrBr) F-395
Br2(cr,l) Bromine (Br_2) F-396
Br2(g) Bromine (Br_2) F-397
Br2Ca1(cr) Calcium Bromide ($CaBr_2$) F-398
Br2Ca1(cr,l) Calcium Bromide ($CaBr_2$) F-399
Br2Ca1(g) Calcium Bromide ($CaBr_2$) F-400
Br2Ca1(l) Calcium Bromide ($CaBr_2$) F-401
Br2Fe1(cr) Iron Bromide ($FeBr_2$) F-402
Br2Fe1(cr,l) Iron Bromide ($FeBr_2$) F-403
Br2Fe1(g) Iron Bromide ($FeBr_2$) F-404
Br2Fe1(l) Iron Bromide ($FeBr_2$) F-405
Br2H2Si1(g) Dibromosilane (SiH_2Br_2) F-406
Br2Hg1(cr) Mercury Bromide ($HgBr_2$) F-407
Br2Hg1(cr,l) Mercury Bromide ($HgBr_2$) F-408
Br2Hg1(g) Mercury Bromide ($HgBr_2$) F-409
Br2Hg1(l) Mercury Bromide ($HgBr_2$) F-410
Br2Hg2(cr) Mercury Bromide (Hg_2Br_2) F-411
Br2K2(g) Potassium Bromide ($(KBr)_2$) F-412
Br2Li2(g) Lithium Bromide ($(LiBr)_2$) F-413
Br2Mg1(cr) Magnesium Bromide ($MgBr_2$) F-414
Br2Mg1(cr,l) Magnesium Bromide ($MgBr_2$) F-415
Br2Mg1(g) Magnesium Bromide ($MgBr_2$) F-416
Br2Mg1(l) Magnesium Bromide ($MgBr_2$) F-417
Br2Mg1+(g) Magnesium Bromide, Ion ($MgBr_2^+$) F-418
Br2Mo1(cr) Molybdenum Bromide ($MoBr_2$) F-419
Br2Mo1(g) Molybdenum Bromide ($MoBr_2$) F-420
Br2Na2(g) Sodium Bromide ($(NaBr)_2$) F-421
Br2Pb1(cr) Lead Bromide ($PbBr_2$) F-422
Br2Pb1(cr,l) Lead Bromide ($PbBr_2$) F-423
Br2Pb1(g) Lead Bromide ($PbBr_2$) F-424
Br2Pb1(l) Lead Bromide ($PbBr_2$) F-425
Br2Si1(g) Dibromosilylene ($SiBr_2$) F-426
Br2Sr1(cr) Strontium Bromide ($SrBr_2$) F-427

Br2Sr1(cr,l)	Strontium Bromide ($SrBr_2$)	F-428
Br2Sr1(g)	Strontium Bromide ($SrBr_2$)	F-429
Br2Sr1(l)	Strontium Bromide ($SrBr_2$)	F-430
Br2Ti1(cr)	Titanium Bromide ($TiBr_2$)	F-431
Br2Ti1(g)	Titanium Bromide ($TiBr_2$)	F-432
Br2Zr1(cr)	Zirconium Bromide ($ZrBr_2$)	F-433
Br2Zr1(cr,l)	Zirconium Bromide ($ZrBr_2$)	F-434
Br2Zr1(g)	Zirconium Bromide ($ZrBr_2$)	F-435
Br2Zr1(l)	Zirconium Bromide ($ZrBr_2$)	F-436
Br3H1Si1(g)	Tribromosilane ($SiHBr_3$)	F-437
Br3Mo1(cr)	Molybdenum Bromide ($MoBr_3$)	F-438
Br3Mo1(g)	Molybdenum Bromide ($MoBr_3$)	F-439
Br3O1P1(g)	Phosphoryl Bromide ($POBr_3$)	F-440
Br3P1(g)	Phosphorus Bromide (PBr_3)	F-441
Br3P1S1(g)	Thiophosphoryl Bromide ($PSBr_3$)	F-442
Br3Si1(g)	Tribromosilyl ($SiBr_3$)	F-443
Br3Ti1(cr)	Titanium Bromide ($TiBr_3$)	F-444
Br3Ti1(g)	Titanium Bromide ($TiBr_3$)	F-445
Br3Zr1(cr)	Zirconium Bromide ($ZrBr_3$)	F-446
Br3Zr1(g)	Zirconium Bromide ($ZrBr_3$)	F-447
Br4Fe2(g)	Iron Bromide ($(FeBr_2)_2$)	F-448
Br4Mg2(g)	Magnesium Bromide ($(MgBr_2)_2$)	F-449
Br4Mo1(cr)	Molybdenum Bromide ($MoBr_4$)	F-450
Br4Mo1(g)	Molybdenum Bromide ($MoBr_4$)	F-451
Br4Pb1(g)	Lead Bromide ($PbBr_4$)	F-452
Br4Si1(g)	Tetrabromosilane ($SiBr_4$)	F-453
Br4Si1(l)	Tetrabromosilane ($SiBr_4$)	F-454
Br4Ti1(cr)	Titanium Bromide ($TiBr_4$)	F-455
Br4Ti1(cr,l)	Titanium Bromide ($TiBr_4$)	F-456
Br4Ti1(g)	Titanium Bromide ($TiBr_4$)	F-457
Br4Ti1(l)	Titanium Bromide ($TiBr_4$)	F-458
Br4Zr1(cr)	Zirconium Bromide ($ZrBr_4$)	F-459
Br4Zr1(g)	Zirconium Bromide ($ZrBr_4$)	F-460
Br5Nb1(cr)	Niobium Bromide ($NbBr_5$)	F-461
Br5Nb1(cr,l)	Niobium Bromide ($NbBr_5$)	F-462
Br5Nb1(g)	Niobium Bromide ($NbBr_5$)	F-463
Br5Nb1(l)	Niobium Bromide ($NbBr_5$)	F-464
Br5W1(cr)	Tungsten Bromide (WBr_5)	F-465
Br5W1(cr,l)	Tungsten Bromide (WBr_5)	F-466
Br5W1(g)	Tungsten Bromide (WBr_5)	F-467
Br5W1(l)	Tungsten Bromide (WBr_5)	F-468
Br6W1(cr)	Tungsten Bromide (WBr_6)	F-469
Br6W1(g)	Tungsten Bromide (WBr_6)	F-470
C0.98Nb1(cr)	Niobium Carbide ($NbC_{0.98}$)	F-471
C1(g)	Carbon (C)	F-472
C1-(g)	Carbon, Ion (C^-)	F-473

C1(ref) Carbon (C) F-474
C1+(g) Carbon, Ion (C^+) F-475
C1Al1(g) Aluminum Carbide (AlC) F-476
C1B1(g) Boron Carbide (BC) F-477
C1B4(cr) Boron Carbide (B_4C) F-478
C1B4(cr,l) Boron Carbide (B_4C) F-479
C1B4(l) Boron Carbide (B_4C) F-480
C1Be2(cr) Beryllium Carbide (Be_2C) F-481
C1Be2(cr,l) Beryllium Carbide (Be_2C) F-482
C1Be2(l) Beryllium Carbide (Be_2C) F-483
C1Br1(g) Bromomethylidyne (CBr) F-484
C1Br1F3(g) Bromotrifluoromethane ($CBrF_3$) F-485
C1Br1N1(g) Cyanogen Bromide (BrCN) F-486
C1Br4(g) Tetrabromomethane (CBr_4) F-487
C1Cl1(g) Chloromethylidyne (CCl) F-488
C1Cl1F1O1(g) Carbonic Chloride Fluoride (COClF) F-489
C1Cl1F3(g) Chlorotrifluoromethane ($CClF_3$) F-490
C1Cl1N1(g) Cyanogen Chloride (ClCN) F-491
C1Cl1O1(g) Carbonyl Chloride (COCl) F-492
C1Cl2(g) Dichloromethylene (CCl_2) F-493
C1Cl2F2(g) Dichlorodifluoromethane (CCl_2F_2) F-494
C1Cl2O1(g) Carbonic Dichloride ($COCl_2$) F-495
C1Cl3(g) Trichloromethyl (CCl_3) F-496
C1Cl3F1(g) Triclorofluormethane (CCl_3F) F-497
C1Cl4(g) Tetrachloromethane (CCl_4) F-498
C1Cu1N1(cr) Copper Cyanide (CuCN) F-499
C1F1(g) Fluoromethylidyne (CF) F-500
C1F1+(g) Fluoromethylidyne, Ion (CF^+) F-501
C1F1N1(g) Cyanogen Fluoride (FCN) F-502
C1F1O1(g) Carbonyl Fluoride (COF) F-503
C1F2(g) Difluoromethylene (CF_2) F-504
C1F2+(g) Difluoromethylene, Ion (CF_2^+) F-505
C1F2O1(g) Carbonic Difluoride (COF_2) F-506
C1F3(g) Trifluoromethyl (CF_3) F-507
C1F3+(g) Trifluoromethyl, Ion (CF_3^+) F-508
C1F3I1(g) Trifluoriodomethane (CIF_3) F-509
C1F4(g) Tetrafluoromethane (CF_4) F-510
C1F4O1(g) Trifluoromethyl Hypofluorite (CF_3OF) F-511
C1F8S1(g) Pentafluoro(Trifluoromethyl)Sulfur (CF_3SF_5) F-512
C1H1(g) Methylidyne (CH) F-513
C1H1+(g) Methylidyne, Ion (CH^+) F-514
C1H1Cl1(g) Chloromethylene (CHCl) F-515
C1H1Cl1F2(g) Chlorodifluoromethane ($CHClF_2$) F-516
C1H1Cl2F1(g) Dichlorofluoromethane ($CHCl_2F$) F-517
C1H1Cl3(g) Trichloromethane ($CHCl_3$) F-518
C1H1F1(g) Fluoromethylene (CHF) F-519

C1H1F1O1(g) Formyl Fluoride (HCOF) F-520
C1H1F3(g) Trifluoromethane (CHF_3) F-521
C1H1N1(g) Hydrogen Cyanide (HCN) F-522
C1H1N1O1(g) Hydrogen Isocyanate (HNCO) F-523
C1H1O1(g) Formyl (HCO) F-524
C1H1O1+(g) Formyl, Ion (HCO^+) F-525
C1H1P1(g) Methinophosphide (CHP) F-526
C1H2(g) Methylene (CH_2) F-527
C1H2Cl1F1(g) Chlorofluoromethane (CH_2ClF) F-528
C1H2Cl2(g) Dichloromethane (CH_2Cl_2) F-529
C1H2F2(g) Difluoromethane (CH_2F_2) F-530
C1H2O1(g) Formaldehyde (H_2CO) F-531
C1H3(g) Methyl (CH_3) F-532
C1H3Cl1(g) Chloromethane (CH_3Cl) F-533
C1H3Cl3Si1(g) Trichloromethylsilane ($SiCH_3Cl_3$) F-534
C1H3F1(g) Fluoromethane (CH_3F) F-535
C1H3F3Si1(g) Trifluoromethylsilane ($SiCH_3F_3$) F-536
C1H4(g) Methane (CH_4) F-537
C1I1N1(g) Cyanogen Iodide (CNI) F-538
C1K1N1(cr) Potassium Cyanide (KCN) F-539
C1K1N1(cr,l) Potassium Cyanide (KCN) F-540
C1K1N1(g) Potassium Cyanide (KCN) F-541
C1K1N1(l) Potassium Cyanide (KCN) F-542
C1K2O3(cr) Potassium Carbonate (K_2CO_3) F-543
C1K2O3(cr,l) Potassium Carbonate (K_2CO_3) F-544
C1K2O3(l) Potassium Carbonate (K_2CO_3) F-545
C1Li2O3(cr) Lithium Carbonate (Li_2CO_3) F-546
C1Li2O3(cr,l) Lithium Carbonate (Li_2CO_3) F-547
C1Li2O3(l) Lithium Carbonate (Li_2CO_3) F-548
C1Mg1O3(cr) Magnesium Carbonate ($MgCO_3$) F-549
C1N1(g) Cyanogen (CN) F-550
C1N1-(g) Cyanide (CN^-) F-551
C1N1+(g) Cyanogen, Ion (CN^+) F-552
C1N1Na1(cr) Sodium Cyanide (NaCN) F-553
C1N1Na1(cr,l) Sodium Cyanide (NaCN) F-554
C1N1Na1(g) Sodium Cyanide (NaCN) F-555
C1N1Na1(l) Sodium Cyanide (NaCN) F-556
C1N1O1(g) NCO Radical (NC_{101}) F-557
C1N2(g)_01 CNN Radical (CNN) F-558
C1N2(g)_02 NCN Radical (NCN) F-559
C1Na2O3(cr) Sodium Carbonate (Na_2CO_3) F-560
C1Na2O3(cr,l) Sodium Carbonate ($NaCO_3$) F-561
C1Na2O3(l) Sodium Carbonate (Na_2CO_3) F-562
C1O1(g) Carbon Monoxide (CO) F-563
C1O1S1(g) Carbon Oxide Sulfide (COS) F-564
C1O2(g) Carbon Dioxide (CO_2) F-565

C1O2-(g)	Carbon Dioxide, Ion (CO_2^-)	F-566
C1P1(g)	Carbon Phosphide (CP)	F-567
C1S1(g)	Carbon Sulfide (CS)	F-568
C1S2(g)	Carbon Disulfide (CS_2)	F-569
C1Si1(cr)_01	Silicon Carbide, Alpha (SiC)	F-570
C1Si1(cr)_02	Silicon Carbide, Beta (SiC)	F-571
C1Si1(g)	Silicon Carbide (SiC)	F-572
C1Si2(g)	Silicon Carbide (Si_2C)	F-573
C1Ta1(cr)	Tantalum Carbide (TaC)	F-574
C1Ta1(cr,l)	Tantalum Carbide (TaC)	F-575
C1Ta1(l)	Tantalum Carbide (TaC)	F-576
C1Ti1(cr)	Titanium Carbide (TiC)	F-577
C1Ti1(cr,l)	Titanium Carbide (TiC)	F-578
C1Ti1(l)	Titanium Carbide (TiC)	F-579
C1Zr1(cr)	Zirconium Carbide (ZrC)	F-580
C1Zr1(cr,l)	Zirconium Carbide (ZrC)	F-581
C1Zr1(l)	Zirconium Carbide (ZrC)	F-582
C2(g)	Carbon (C_2)	F-583
C2-(g)	Carbon, Ion (C_2^-)	F-584
C2Be1(g)	Beryllium Carbide (BeC_2)	F-585
C2Cl2(g)	Dichloroethyne (C_2Cl_2)	F-586
C2Cl4(g)	Tetrachloroethene (C_2Cl_4)	F-587
C2Cl6(g)	Hexachloroethane (C_2Cl_6)	F-588
C2Cr3(cr)	Chromium Carbide (Cr_3C_2)	F-589
C2F2(g)	Difluoroethyne (C_2F_2)	F-590
C2F3N1(g)	Trifluoroacetonitrile (CF_3CN)	F-591
C2F4(g)	Tetrafluoroethene (C_2F_4)	F-592
C2F6(g)	Hexafluoroethane (C_2F_6)	F-593
C2H1(g)	Ethynyl (C_2H)	F-594
C2H1Cl1(g)	Chloroethyne (C_2HCl)	F-595
C2H1F1(g)	Fluoroethyne (C_2HF)	F-596
C2H2(g)	Ethyne (C_2H_2)	F-597
C2H4(g)	Ethene (C_2H_4)	F-598
C2H4O1(g)	Oxirane (C_2H_4O)	F-599
C2K2N2(g)	Potassium Cyanide ($(KCN)_2$)	F-600
C2Li2(cr)	Lithium Carbide (Li_2C_2)	F-601
C2Mg1(cr)	Magnesium Carbide (MgC_2)	F-602
C2N1(g)	CNC Radical (C_2N)	F-603
C2N2(g)	Ethanedinitrile ($(CN)_2$)	F-604
C2N2Na2(g)	Sodium Cyanide ($(NaCN)_2$)	F-605
C2O1(g)	CCO Radical (C_2O)	F-606
C2Si1(g)	Silicon Carbide (SiC_2)	F-607
C3(g)	Carbon (C_3)	F-608
C3Al4(cr)	Aluminum Carbide (Al_4C_3)	F-609
C3Cr7(cr)	Chromium Carbide (Cr_7C_3)	F-610
C3Mg2(cr)	Magnesium Carbide (Mg_2C_3)	F-611

C3O2(g)	Carbon Suboxide (C_3O_2)	F-612
C4(g)	Carbon (C_4)	F-613
C4H12Si1(g)	Tetramethylsilane ($Si(CH_3)_4$)	F-614
C4N2(g)	$_2$-Butynedinitrile (C_4N_2)	F-615
C4Ni1O4(g)	Nickel Carbonyl ($Ni(CO)_4$)	F-616
C4Ni1O4(l)	Nickel Carbonyl ($Ni(CO)_4$)	F-617
C5(g)	Carbon (C_5)	F-618
C5Fe1O5(g)	Iron Carbonyl ($Fe(CO)_5$)	F-619
C5Fe1O5(l)	Iron Carbonyl ($Fe(CO)_5$)	F-620
C6Cr23(cr)	Chromium Carbide ($Cr_{23}C_6$)	F-621
Ca1(cr)_01	Calcium, Alpha (Ca)	F-622
Ca1(cr)_02	Calcium, Beta (Ca)	F-623
Ca1(cr,l)	Calcium (Ca)	F-624
Ca1(g)	Calcium (Ca)	F-625
Ca1(l)	Calcium (Ca)	F-626
Ca1+(g)	Calcium, Ion (Ca^+)	F-627
Ca1Cl1(g)	Calcium Chloride (CaCl)	F-628
Ca1Cl2(cr)	Calcium Chloride ($CaCl_2$)	F-629
Ca1Cl2(cr,l)	Calcium Chloride ($CaCl_2$)	F-630
Ca1Cl2(g)	Calcium Chloride ($CaCl_2$)	F-631
Ca1Cl2(l)	Calcium Chloride ($CaCl_2$)	F-632
Ca1F1(g)	Calcium Fluoride (CaF)	F-633
Ca1F2(cr)	Calcium Fluoride (CaF_2)	F-634
Ca1F2(cr,l)	Calcium Fluoride (CaF_2)	F-635
Ca1F2(g)	Calcium Fluoride (CaF_2)	F-636
Ca1F2(l)	Calcium Fluoride (CaF_2)	F-637
Ca1H1O1(g)	Calcium Hydroxide (CaOH)	F-638
Ca1H1O1+(g)	Calcium Hydroxide, Ion ($CaOH^+$)	F-639
Ca1H2O2(cr)	Calcium Hydroxide ($Ca(OH)_2$)	F-640
Ca1H2O2(g)	Calcium Hydroxide ($Ca(OH)_2$)	F-641
Ca1I1(g)	Calcium Iodide (CaI)	F-642
Ca1I2(cr)	Calcium Iodide (CaI_2)	F-643
Ca1I2(cr,l)	Calcium Iodide (CaI_2)	F-644
Ca1I2(g)	Calcium Iodide (CaI_2)	F-645
Ca1I2(l)	Calcium Iodide (CaI_2)	F-646
Ca1O1(cr)	Calcium Oxide (CaO)	F-647
Ca1O1(cr,l)	Calcium Oxide (CaO)	F-648
Ca1O1(g)	Calcium Oxide (CaO)	F-649
Ca1O1(l)	Calcium Oxide (CaO)	F-650
Ca1S1(cr)	Calcium Sulfide (CaS)	F-651
Ca1S1(g)	Calcium Sulfide (CaS)	F-652
Ca2(g)	Calcium (Ca_2)	F-653
Cl1(g)	Chlorine (Cl)	F-654
Cl1-(g)	Chlorine, Ion (Cl^-)	F-655
Cl1+(g)	Chlorine, Ion (Cl^+)	F-656
Cl10W2(g)	Tungsten Chloride ($(WCl_5)_2$)	F-657

Cl1Co1(g)	Cobalt Chloride (CoCl)	F-658
Cl1Cs1(cr)	Cesium Chloride (CsCl)	F-659
Cl1Cs1(cr,l)	Cesium Chloride (CsCl)	F-660
Cl1Cs1(g)	Cesium Chloride (CsCl)	F-661
Cl1Cs1(l)	Cesium Chloride (CsCl)	F-662
Cl1Cu1(cr)	Copper Chloride (CuCl)	F-663
Cl1Cu1(cr,l)	Copper Chloride (CuCl)	F-664
Cl1Cu1(g)	Copper Chloride (CuCl)	F-665
Cl1Cu1(l)	Copper Chloride (CuCl)	F-666
Cl1D1(g)	Hydrochloric Acid-D (DCl)	F-667
Cl1D1O1(g)	Hypochlorous Acid-D (DOCl)	F-668
Cl1F1(g)	Chlorine Fluoride (ClF)	F-669
Cl1F1Li2(g)	Lithium Chloride Fluoride (Li_2ClF)	F-670
Cl1F1Mg1(g)	Magnesium Chloride Fluoride (MgClF)	F-671
Cl1F1O2S1(g)	Sulfuryl Chloride Fluoride (SO_2ClF)	F-672
Cl1F1O3(g)	Perchloryl Fluoride (ClO_3F)	F-673
Cl1F2O1P1(g)	Phosphoryl Chloride Fluoride ($POClF_2$)	F-674
Cl1F3(g)	Chlorine Fluoride (ClF_3)	F-675
Cl1F3Si1(g)	Chlorotrifluorosilane ($SiClF_3$)	F-676
Cl1F5(g)	Chlorine Fluoride (ClF_5)	F-677
Cl1F5S1(g)	Sulfur Chloride Fluoride ($SClF_5$)	F-678
Cl1Fe1(g)	Iron Chloride (FeCl)	F-679
Cl1H1(g)	Hydrogen Chloride (HCl)	F-680
Cl1H1O1(g)	Hypochlorous Acid (HOCl)	F-681
Cl1H3Si1(g)	Chlorosilane (SiH_3Cl)	F-682
Cl1H4N1(cr)	Ammonium Chloride (NH_4Cl)	F-683
Cl1H4N1O4(cr)	Ammonium Perchlorate (NH_4ClO_4)	F-684
Cl1Hg1(g)	Mercury Chloride (HgCl)	F-685
Cl1I1(cr)	Iodine Chloride (ICl)	F-686
Cl1I1(cr,l)	Iodine Chloride (ICl)	F-687
Cl1I1(g)	Iodine Chloride (ICl)	F-688
Cl1I1(l)	Iodine Chloride (ICl)	F-689
Cl1K1(cr)	Potassium Chloride (KCl)	F-690
Cl1K1(cr,l)	Potassium Chloride (KCl)	F-691
Cl1K1(g)	Potassium Chloride (KCl)	F-692
Cl1K1(l)	Potassium Chloride (KCl)	F-693
Cl1K1O4(cr)	Potassium Perchlorate ($KClO_4$)	F-694
Cl1Li1(cr)	Lithium Chloride (LiCl)	F-695
Cl1Li1(cr,l)	Lithium Chloride (LiCl)	F-696
Cl1Li1(g)	Lithium Chloride (LiCl)	F-697
Cl1Li1(l)	Lithium Chloride (LiCl)	F-698
Cl1Li1O1(g)	Lithium Hypochlorite (LiOCl)	F-699
Cl1Li1O4(cr)	Lithium Perchlorate ($LiClO_4$)	F-700
Cl1Li1O4(cr,l)	Lithium Perchlorate ($LiClO_4$)	F-701
Cl1Li1O4(l)	Lithium Perchlorate ($LiClO_4$)	F-702
Cl1Mg1(g)	Magnesium Chloride (MgCl)	F-703

Cl1Mg1+(g)	Magnesium Chloride, Ion ($MgCl^+$)	F-704
Cl1N1O1(g)	Nitrosyl Chloride (ONCl)	F-705
Cl1N1O2(g)	Nitryl Chloride (NO_2Cl)	F-706
Cl1Na1(cr)	Sodium Chloride (NaCl)	F-707
Cl1Na1(cr,l)	Sodium Chloride (NaCl)	F-708
Cl1Na1(g)	Sodium Chloride (NaCl)	F-709
Cl1Na1(l)	Sodium Chloride (NaCl)	F-710
Cl1Na1O4(cr)	Sodium Perchlorate ($NaClO_4$)	F-711
Cl1Ni1(g)	Nickel Chloride (NiCl)	F-712
Cl1O1(g)	Chlorine Oxide (ClO)	F-713
Cl1O1Ti1(g)	Titanium Chloride Oxide (OTiCl)	F-714
Cl1O2(g)	Chlorine Oxide (ClO_2)	F-715
Cl1P1(g)	Phosphorus Chloride (PCl)	F-716
Cl1Pb1(g)	Lead Chloride (PbCl)	F-717
Cl1Pb1+(g)	Lead Chloride, Ion ($PbCl^+$)	F-718
Cl1S1(g)	Sulfur Chloride (SCl)	F-719
Cl1S1+(g)	Sulfur Chloride, Ion (SCl^+)	F-720
Cl1S2(g)	Sulfur Chloride (S_2Cl)	F-721
Cl1Si1(g)	Clorosilylidyne (SiCl)	F-722
Cl1Sr1(g)	Strontium Chloride (SrCl)	F-723
Cl1Ti1(g)	Titanium Chloride (TiCl)	F-724
Cl1W1(g)	Tungsten Chloride (WCl)	F-725
Cl1Zr1(g)	Zirconium Chloride (ZrCl)	F-726
Cl2(ref)	Chlorine (Cl_2)	F-727
Cl2Co1(cr)	Cobalt Chloride ($CoCl_2$)	F-728
Cl2Co1(cr,l)	Cobalt Chloride ($CoCl_2$)	F-729
Cl2Co1(g)	Cobalt Chloride ($CoCl_2$)	F-730
Cl2Co1(l)	Cobalt Chloride ($CoCl_2$)	F-731
Cl2Cs2(g)	Cesium Chloride ($(CsCl)_2$)	F-732
Cl2Cu1(cr)	Copper Chloride ($CuCl_2$)	F-733
Cl2F1O1P1(g)	Phosphoryl Chloride Fluoride ($POCl_2F$)	F-734
Cl2Fe1(cr)	Iron Chloride ($FeCl_2$)	F-735
Cl2Fe1(cr,l)	Iron Chloride ($FeCl_2$)	F-736
Cl2Fe1(g)	Iron Chloride ($FeCl_2$)	F-737
Cl2Fe1(l)	Iron Chloride ($FeCl_2$)	F-738
Cl2H2Si1(g)	Dichlorosilane (SiH_2Cl_2)	F-739
Cl2Hg1(cr)	Mercury Chloride ($HgCl_2$)	F-740
Cl2Hg1(cr,l)	Mercury Chloride ($HgCl_2$)	F-741
Cl2Hg1(g)	Mercury Chloride ($HgCl_2$)	F-742
Cl2Hg1(l)	Mercury Chloride ($HgCl_2$)	F-743
Cl2Hg2(cr)	Mercury Chloride (Hg_2Cl_2)	F-744
Cl2K2(g)	Potassium Chloride ($(KCl)_2$)	F-745
Cl2Li2(g)	Lithium Chloride ($(LiCl)_2$)	F-746
Cl2Mg1(cr)	Magnesium Chloride ($MgCl_2$)	F-747
Cl2Mg1(cr,l)	Magnesium Chloride ($MgCl_2$)	F-748
Cl2Mg1(g)	Magnesium Chloride ($MgCl_2$)	F-749

Cl2Mg1(l)	Magnesium Chloride ($MgCl_2$)	F-750
Cl2Mo1O2(g)	Molybdenum Chloride Oxide (MoO_2Cl_2)	F-751
Cl2Na2(g)	Sodium Chloride ($(NaCl)_2$)	F-752
Cl2Ni1(cr)	Nickel Chloride ($NiCl_2$)	F-753
Cl2Ni1(cr,l)	Nickel Chloride ($NiCl_2$)	F-754
Cl2Ni1(g)	Nickel Chloride ($NiCl_2$)	F-755
Cl2Ni1(l)	Nickel Chloride ($NiCl_2$)	F-756
Cl2O1(g)	Chlorine Oxide (Cl_2O)	F-757
Cl2O1Ti1(g)	Titanium Chloride Oxide ($TiOCl_2$)	F-758
Cl2O2S1(g)	Sulfuryl Chloride (SO_2Cl_2)	F-759
Cl2O2W1(cr)	Tungsten Chloride Oxide (WO_2Cl_2)	F-760
Cl2O2W1(g)	Tungsten Chloride Oxide (WO_2Cl_2)	F-761
Cl2Pb1(cr)	Lead Chloride ($PbCl_2$)	F-762
Cl2Pb1(cr,l)	Lead Chloride ($PbCl_2$)	F-763
Cl2Pb1(g)	Lead Chloride ($PbCl_2$)	F-764
Cl2Pb1(l)	Lead Chloride ($PbCl_2$)	F-765
Cl2Pb1+(g)	Lead Chloride, Ion ($PbCl_2^+$)	F-766
Cl2S1(g)	Sulfur Chloride (SCl_2)	F-767
Cl2S1(l)	Sulfur Chloride (SCl_2)	F-768
Cl2S1+(g)	Sulfur Chloride, Ion (SCl_2^+)	F-769
Cl2S2(g)	Sulfur Chloride (ClSSCl)	F-770
Cl2S2(l)	Sulfur Chloride (ClSSCl)	F-771
Cl2Si1(g)	Dichlorosilylene ($SiCl_2$)	F-772
Cl2Sr1(cr)	Strontium Chloride ($SrCl_2$)	F-773
Cl2Sr1(cr,l)	Strontium Chloride ($SrCl_2$)	F-774
Cl2Sr1(g)	Strontium Chloride ($SrCl_2$)	F-775
Cl2Sr1(l)	Strontium Chloride ($SrCl_2$)	F-776
Cl2Ti1(cr)	Titanium Chloride ($TiCl_2$)	F-777
Cl2Ti1(g)	Titanium Chloride ($TiCl_2$)	F-778
Cl2W1(cr)	Tungsten Chloride (WCl_2)	F-779
Cl2W1(g)	Tungsten Chloride (WCl_2)	F-780
Cl2Zr1(cr)	Zirconium Chloride ($ZrCl_2$)	F-781
Cl2Zr1(cr,l)	Zirconium Chloride ($ZrCl_2$)	F-782
Cl2Zr1(g)	Zirconium Chloride ($ZrCl_2$)	F-783
Cl2Zr1(l)	Zirconium Chloride ($ZrCl_2$)	F-784
Cl3Co1(g)	Cobalt Chloride ($CoCl_3$)	F-785
Cl3Cu3(g)	Copper Chloride ($(CuCl)_3$)	F-786
Cl3F1Si1(g)	Trichlorofluorosilane ($SiCl_3F$)	F-787
Cl3Fe1(cr)	Iron Chloride ($FeCl_3$)	F-788
Cl3Fe1(cr,l)	Iron Chloride ($FeCl_3$)	F-789
Cl3Fe1(g)	Iron Chloride ($FeCl_3$)	F-790
Cl3Fe1(l)	Iron Chloride ($FeCl_3$)	F-791
Cl3H1Si1(g)	Trichlorosilane ($SiHCl_3$)	F-792
Cl3Li3(g)	Lithium Chloride ($(LiCl)_3$)	F-793
Cl3O1P1(g)	Phosphoryl Chloride ($OPCl_3$)	F-794
Cl3P1(g)	Phosphorus Chloride (PCl_3)	F-795

Cl3P1S1(g)	Thiophosphoryl Chloride ($SPCl_3$)	F-796
Cl3Si1(g)	Trichlorosilyl ($SiCl_3$)	F-797
Cl3Ti1(cr)	Titanium Chloride ($TiCl_3$)	F-798
Cl3Ti1(g)	Titanium Chloride ($TiCl_3$)	F-799
Cl3Zr1(cr)	Zirconium Chloride ($ZrCl_3$)	F-800
Cl3Zr1(g)	Zirconium Chloride ($ZrCl_3$)	F-801
Cl4Co2(g)	Cobalt Chloride ($(CoCl_2)_2$)	F-802
Cl4Fe2(g)	Iron Chloride ($(FeCl_2)_2$)	F-803
Cl4Mg2(g)	Magnesium Chloride ($(MgCl_2)_2$)	F-804
Cl4Mo1(cr)	Molybdenum Chloride ($MoCl_4$)	F-805
Cl4Mo1(cr,l)	Molybdenum Chloride ($MoCl_4$)	F-806
Cl4Mo1(g)	Molybdenum Chloride ($MoCl_4$)	F-807
Cl4Mo1(l)	Molybdenum Chloride ($MoCl_4$)	F-808
Cl4O1W1(cr)	Tungsten Chloride Oxide ($OWCl_4$)	F-809
Cl4O1W1(cr,l)	Tungsten Chloride Oxide ($OWCl_4$)	F-810
Cl4O1W1(g)	Tungsten Chloride Oxide ($OWCl_4$)	F-811
Cl4O1W1(l)	Tungsten Chloride Oxide ($OWCl_4$)	F-812
Cl4Pb1(g)	Lead Chloride ($PbCl_4$)	F-813
Cl4Si1(g)	Tetrachlorosilane ($SiCl_4$)	F-814
Cl4Ti1(cr)	Titanium Chloride ($TiCl_4$)	F-815
Cl4Ti1(cr,l)	Titanium Chloride ($TiCl_4$)	F-816
Cl4Ti1(g)	Titanium Chloride ($TiCl_4$)	F-817
Cl4Ti1(l)	Titanium Chloride ($TiCl_4$)	F-818
Cl4W1(cr)	Tungsten Chloride (WCl_4)	F-819
Cl4W1(g)	Tungsten Chloride (WCl_4)	F-820
Cl4Zr1(cr)	Zirconium Chloride ($ZrCl_4$)	F-821
Cl4Zr1(g)	Zirconium Chloride ($ZrCl_4$)	F-822
Cl5Mo1(cr)	Molybdenum Chloride ($MoCl_5$)	F-823
Cl5Mo1(cr,l)	Molybdenum Chloride ($MoCl_5$)	F-824
Cl5Mo1(g)	Molybdenum Chloride ($MoCl_5$)	F-825
Cl5Mo1(l)	Molybdenum Chloride ($MoCl_5$)	F-826
Cl5Nb1(cr)	Niobium Chloride ($NbCl_5$)	F-827
Cl5Nb1(cr,l)	Niobium Chloride ($NbCl_5$)	F-828
Cl5Nb1(g)	Niobium Chloride ($NbCl_5$)	F-829
Cl5Nb1(l)	Niobium Chloride ($NbCl_5$)	F-830
Cl5P1(g)	Phosphorus Chloride (PCl_5)	F-831
Cl5Ta1(cr)	Tantalum Chloride ($TaCl_5$)	F-832
Cl5Ta1(cr,l)	Tantalum Chloride ($TaCl_5$)	F-833
Cl5Ta1(g)	Tantalum Chloride ($TaCl_5$)	F-834
Cl5Ta1(l)	Tantalum Chloride ($TaCl_5$)	F-835
Cl5W1(cr)	Tungsten Chloride (WCl_5)	F-836
Cl5W1(cr,l)	Tungsten Chloride (WCl_5)	F-837
Cl5W1(g)	Tungsten Chloride (WCl_5)	F-838
Cl5W1(l)	Tungsten Chloride (WCl_5)	F-839
Cl6Fe2(g)	Iron Chloride ($(FeCl_3)_2$)	F-840
Cl6Mo1(cr)	Molybdenum Chloride ($MoCl_6$)	F-841

Cl6Mo1(g)	Molybdenum Chloride ($MoCl_6$)	F-842
Cl6W1(cr)_01	Tungsten Chloride, Alpha (WCl_6)	F-843
Cl6W1(cr)_02	Tungsten Chloride, Beta (WCl_6)	F-844
Cl6W1(cr,l)	Tungsten Chloride (WCl_6)	F-845
Cl6W1(g)	Tungsten Chloride (WCl_6)	F-846
Cl6W1(l)	Tungsten Chloride (WCl_6)	F-847
Co1(cr)	Cobalt (Co)	F-848
Co1(cr,l)	Cobalt (Co)	F-849
Co1(g)	Cobalt (Co)	F-850
Co1-(g)	Cobalt, Ion (Co^-)	F-851
Co1(l)	Cobalt (Co)	F-852
Co1+(g)	Cobalt, Ion (Co^+)	F-853
Co1F2(cr)	Cobalt Fluoride (CoF_2)	F-854
Co1F2(cr,l)	Cobalt Fluoride (CoF_2)	F-855
Co1F2(g)	Cobalt Fluoride (CoF_2)	F-856
Co1F2(l)	Cobalt Fluoride (CoF_2)	F-857
Co1F3(cr)	Cobalt Fluoride (CoF_3)	F-858
Co1O1(cr)	Cobalt Oxide (CoO)	F-859
Co1O4S1(cr)	Cobalt Sulfate ($CoSO_4$)	F-860
Co3O4(cr)	Cobalt Oxide (Co_3O_4)	F-861
Cr1(cr)	Chromium (Cr)	F-862
Cr1(cr,l)	Chromium (Cr)	F-863
Cr1(g)	Chromium (Cr)	F-864
Cr1-(g)	Chromium, Ion (Cr^-)	F-865
Cr1(l)	Chromium (Cr)	F-866
Cr1+(g)	Chromium, Ion (Cr^+)	F-867
Cr1N1(cr)	Chromium Nitride (CrN)	F-868
Cr1N1(g)	Chromium Nitride (CrN)	F-869
Cr1O1(g)	Chromium Oxide (CrO)	F-870
Cr1O2(g)	Chromium Oxide (CrO_2)	F-871
Cr1O3(g)	Chromium Oxide (CrO_3)	F-872
Cr2N1(cr)	Chromium Nitride (Cr_2N)	F-873
Cr2O3(cr)	Chromium Oxide (Cr_2O_3)	F-874
Cr2O3(cr,l)	Chromium Oxide (Cr_2O_3)	F-875
Cr2O3(l)	Chromium Oxide (Cr_2O_3)	F-876
Cs1(cr)	Cesium (Cs)	F-877
Cs1(cr,l)	Cesium (Cs)	F-878
Cs1(g)	Cesium (Cs)	F-879
Cs1-(g)	Cesium, Ion (Cs^-)	F-880
Cs1(l)	Cesium (Cs)	F-881
Cs1+(g)	Cesium, Ion (Cs^+)	F-882
Cs1F1(cr)	Cesium Fluoride (CsF)	F-883
Cs1F1(cr,l)	Cesium Fluoride (CsF)	F-884
Cs1F1(g)	Cesium Fluoride (CsF)	F-885
Cs1F1(l)	Cesium Fluoride (CsF)	F-886
Cs1H1O1(cr)	Cesium Hydroxide (CsOH)	F-887

Cs1H1O1(cr,l)	Cesium Hydroxide (CsOH)	F-888
Cs1H1O1(g)	Cesium Hydroxide (CsOH)	F-889
Cs1H1O1(l)	Cesium Hydroxide (CsOH)	F-890
Cs1H1O1+(g)	Cesium Hydroxide, Ion ($CsOH^+$)	F-891
Cs1O1(g)	Cesium Oxide (CsO)	F-892
Cs2(g)	Cesium (Cs_2)	F-893
Cs2F2(g)	Cesium Fluoride ($(CsF)_2$)	F-894
Cs2H2O2(g)	Cesium Hydroxide ($(CsOH)_2$)	F-895
Cs2O1(g)	Cesium Oxide (Cs_2O)	F-896
Cs2O4S1(cr)_01	Cesium Sulfate, I (Cs_2SO_4)	F-897
Cs2O4S1(cr)_02	Cesium Sulfate, II (Cs_2SO_4)	F-898
Cs2O4S1(cr,l)	Cesium Sulfate (Cs_2SO_4)	F-899
Cs2O4S1(g)	Cesium Sulfate (Cs_2SO_4)	F-900
Cs2O4S1(l)	Cesium Sulfate (Cs_2SO_4)	F-901
Cu1(cr)	Copper (Cu)	F-902
Cu1(cr,l)	Copper (Cu)	F-903
Cu1(g)	Copper (Cu)	F-904
Cu1-(g)	Copper, Ion (Cu^-)	F-905
Cu1(l)	Copper (Cu)	F-906
Cu1+(g)	Copper, Ion (Cu^+)	F-907
Cu1F1(cr)	Copper Fluoride (CuF)	F-908
Cu1F1(g)	Copper Fluoride (CuF)	F-909
Cu1F2(cr)	Copper Fluoride (CuF_2)	F-910
Cu1F2(cr,l)	Copper Fluoride (CuF_2)	F-911
Cu1F2(g)	Copper Fluoride (CuF_2)	F-912
Cu1F2(l)	Copper Fluoride (CuF_2)	F-913
Cu1H2O2(cr)	Copper Hydroxide ($Cu(OH)_2$)	F-914
Cu1O1(cr)	Copper Oxide (CuO)	F-915
Cu1O1(g)	Copper Oxide (CuO)	F-916
Cu1O4S1(cr)	Copper Sulfate ($CuSO_4$)	F-917
Cu2(g)	Copper (Cu_2)	F-918
Cu2O1(cr)	Copper Oxide (Cu_2O)	F-919
Cu2O1(cr,l)	Copper Oxide (Cu_2O)	F-920
Cu2O1(l)	Copper Oxide (Cu_2O)	F-921
Cu2O5S1(cr)	Copper Oxide Sulfate ($CuO.CuSO_4$)	F-922
D1(g)	Deuterium (D)	F-923
D1-(g)	Deuterium, Ion (D^-)	F-924
D1+(g)	Deuterium, Ion (D^+)	F-925
D1F1(g)	Hydrofluoric Acid-D (DF)	F-926
D1H1(g)	Hydrogen-D_1 (HD)	F-927
D1H1-(g)	Hydrogen-D_1, Ion (HD^-)	F-928
D1H1+(g)	Hydrogen-D_1, Ion (HD^+)	F-929
D1H1O1(g)	Water-D_1 (HDO)	F-930
D1N1(g)	Imidogen-D_1 (ND)	F-931
D1O1(g)	Hydroxyl-D_1 (OD)	F-932
D1S1(g)	Mercapto-D_1 (SD)	F-933

D2-(g)	Deuterium, Ion (D_2^-)	F-934
D2(ref)	Deuterium (D_2)	F-935
D2+(g)	Deuterium, Ion (D_2^+)	F-936
D2N1(g)	Amidogen-D_2 (ND_2)	F-937
D2N2(g)	Diazene-D_2, Cis (DNND)	F-938
D2O1(g)	Water-D_2 (D_2O)	F-939
D2S1(g)	Hydrogen Sulfide-D_2 (D_2S)	F-940
D3N1(g)	Ammonia-D_3 (ND_3)	F-941
e1-(ref)	Electron Gas (e^-)	F-942
F1(g)	Fluorine (F)	F-943
F1-(g)	Fluorine, Ion (F^-)	F-944
F1+(g)	Fluorine, Ion (F^+)	F-945
F10Mo2(g)	Molybdenum Fluoride ($(MoF_5)_2$)	F-946
F10S2(g)	Sulfur Fluoride (S_2F_{10})	F-947
F15Mo3(g)	Molybdenum Fluoride ($(MoF_5)_3$)	F-948
F1Fe1(g)	Iron Fluoride (FeF)	F-949
F1H1(g)	Hydrogen Fluoride (HF)	F-950
F1H1O1(g)	Hypofluorous Acid (HOF)	F-951
F1H1O3S1(g)	Fluorosulfuric Acid (HSO_3F)	F-952
F1H3Si1(g)	Fluorosilane (SiH_3F)	F-953
F1Hg1(g)	Mercury Fluoride (HgF)	F-954
F1I1(g)	Iodine Fluoride (IF)	F-955
F1K1(cr)	Potassium Fluoride (KF)	F-956
F1K1(cr,l)	Potassium Fluoride (KF)	F-957
F1K1(g)	Potassium Fluoride (KF)	F-958
F1K1(l)	Potassium Fluoride (KF)	F-959
F1Li1(cr)	Lithium Fluoride (LiF)	F-960
F1Li1(cr,l)	Lithium Fluoride (LiF)	F-961
F1Li1(g)	Lithium Fluoride (LiF)	F-962
F1Li1(l)	Lithium Fluoride (LiF)	F-963
F1Li1O1(g)	Lithium Hypofluorite (LiOF)	F-964
F1Mg1(g)	Magnesium Fluoride (MgF)	F-965
F1Mg1+(g)	Magnesium Fluoride, Ion (MgF^+)	F-966
F1Mo1(g)	Molybdenum Fluoride (MoF)	F-967
F1N1(g)	Fluoroimidogen (NF)	F-968
F1N1O1(g)	Nitrosyl Fluoride (ONF)	F-969
F1N1O2(g)	Nitryl Fluoride (NO_2F)	F-970
F1N1O3(g)	Fluorine Nitrate ($FONO_2$)	F-971
F1Na1(cr)	Sodium Fluoride (NaF)	F-972
F1Na1(cr,l)	Sodium Fluoride (NaF)	F-973
F1Na1(g)	Sodium Fluoride (NaF)	F-974
F1Na1(l)	Sodium Fluoride (NaF)	F-975
F1O1(g)	Oxygen Fluoride (OF)	F-976
F1O1Ti1(g)	Titanium Fluoride Oxide (OTiF)	F-977
F1O2(g)	Oxygen Fluoride (O_2F)	F-978
F1P1(g)	Phosphorus Fluoride (PF)	F-979

F1P1-(g) Phosphorus Fluoride, Ion (PF^-) F-980
F1P1+(g) Phosphorus Fluoride, Ion (PF^+) F-981
F1P1S1(g) Thiophosphoryl Fluoride (PSF) F-982
F1Pb1(g) Lead Fluoride (PbF) F-983
F1S1(g) Sulfur Fluoride (SF) F-984
F1S1-(g) Sulfur Fluoride, Ion (SF^-) F-985
F1S1+(g) Sulfur Fluoride, Ion (SF^+) F-986
F1Si1(g) Fluorosilylidyne (SiF) F-987
F1Sr1(g) Strontium Fluoride (SrF) F-988
F1Sr1+(g) Strontium Fluoride, Ion (SrF^+) F-989
F1Ti1(g) Titanium Fluoride (TiF) F-990
F1W1(g) Tungsten Fluoride (WF) F-991
F1Zr1(g) Zirconium Fluoride (ZrF) F-992
F2(ref) Fluorine (F_2) F-993
F2Fe1(cr) Iron Fluoride (FeF_2) F-994
F2Fe1(cr,l) Iron Fluoride (FeF_2) F-995
F2Fe1(g) Iron Fluoride (FeF_2) F-996
F2Fe1(l) Iron Fluoride (FeF_2) F-997
F2H1K1(cr) Potassium Fluoride ($K(HF_2)$) F-998
F2H1K1(cr,l) Potassium Fluoride ($K(HF_2)$) F-999
F2H1K1(l) Potassium Fluoride ($K(HF_2)$) F-1000
F2H2(g) Hydrogen Fluoride ($(HF)_2$) F-1001
F2H2Si1(g) Difluorosilane (SiH_2F_2) F-1002
F2Hg1(cr) Mercury Fluoride (HgF_2) F-1003
F2Hg1(cr,l) Mercury Fluoride (HgF_2) F-1004
F2Hg1(g) Mercury Fluoride (HgF_2) F-1005
F2Hg1(l) Mercury Fluoride (HgF_2) F-1006
F2Hg2(cr) Mercury Fluoride (Hg_2F_2) F-1007
F2K1-(g) Potassium Fluoride, Ion (KF_2^-) F-1008
F2K2(g) Potassium Fluoride ($(KF)_2$) F-1009
F2Li1-(g) Lithium Fluoride, Ion (LiF_2^-) F-1010
F2Li2(g) Lithium Fluoride ($(LiF)_2$) F-1011
F2Mg1(cr) Magnesium Fluoride (MgF_2) F-1012
F2Mg1(cr,l) Magnesium Fluoride (MgF_2) F-1013
F2Mg1(g) Magnesium Fluoride (MgF_2) F-1014
F2Mg1(l) Magnesium Fluoride (MgF_2) F-1015
F2Mg1+(g) Magnesium Fluoride, Ion (MgF_2^+) F-1016
F2Mo1(g) Molybdenum Fluoride (MoF_2) F-1017
F2N1(g) Difluoroamidogen (NF_2) F-1018
F2N2(g)_01 Nitrogen Fluoride, Cis (FNNF) F-1019
F2N2(g)_02 Nitrogen Fluoride, Trans (FNNF) F-1020
F2Na1-(g) Sodium Fluoride, Ion (NaF_2^-) F-1021
F2Na2(g) Sodium Fluoride ($(NaF)_2$) F-1022
F2O1(g) Oxygen Fluoride (OF_2) F-1023
F2O1S1(g) Thionyl Fluoride (OSF_2) F-1024
F2O1Si1(g) Difluorooxosilane ($OSiF_2$) F-1025

F2O1Ti1(g) Titanium Fluoride Oxide ($OTiF_2$) F-1026
F2O2S1(g) Sulfuryl Fluoride (SO_2F_2) F-1027
F2P1(g) Phosphorus Fluoride (PF_2) F-1028
F2P1-(g) Phosphorus Fluoride, Ion (PF_2^-) F-1029
F2P1+(g) Phosphorus Fluoride, Ion (PF_2^+) F-1030
F2Pb1(cr)_01 Lead Fluoride, Alpha (PbF_2) F-1031
F2Pb1(cr)_02 Lead Fluoride, Beta (PbF_2) F-1032
F2Pb1(cr,l) Lead Fluoride (PbF_2) F-1033
F2Pb1(g) Lead Fluoride (PbF_2) F-1034
F2Pb1(l) Lead Fluoride (PbF_2) F-1035
F2S1(g) Sulfur Fluoride (SF_2) F-1036
F2S1-(g) Sulfur Fluoride, Ion (SF_2^-) F-1037
F2S1+(g) Sulfur Fluoride, Ion (SF_2^+) F-1038
F2S2(g)_01 Difluorodisulfane (FSSF) F-1039
F2S2(g)_02 Thiothionyl Fluoride (SSF_2) F-1040
F2Si1(g) Difluorosilylene (SiF_2) F-1041
F2Sr1(cr) Strontium Fluoride (SrF_2) F-1042
F2Sr1(cr,l) Strontium Fluoride (SrF_2) F-1043
F2Sr1(g) Strontium Fluoride (SrF_2) F-1044
F2Sr1(l) Strontium Fluoride (SrF_2) F-1045
F2Ti1(g) Titanium Fluoride (TiF_2) F-1046
F2Zr1(cr) Zirconium Fluoride (ZrF_2) F-1047
F2Zr1(cr,l) Zirconium Fluoride (ZrF_2) F-1048
F2Zr1(g) Zirconium Fluoride (ZrF_2) F-1049
F2Zr1(l) Zirconium Fluoride (ZrF_2) F-1050
F3Fe1(cr) Iron Fluoride (FeF_3) F-1051
F3Fe1(g) Iron Fluoride (FeF_3) F-1052
F3H1Si1(g) Trifluorosilane ($SiHF_3$) F-1053
F3H3(g) Hydrogen Fluoride (H_3F_3) F-1054
F3Li3(g) Lithium Fluoride ($(LiF)_3$) F-1055
F3Mo1(g) Molybdenum Fluoride (MoF_3) F-1056
F3N1(g) Nitrogen Fluoride (NF_3) F-1057
F3N1O1(g) Nitrogen Fluoride Oxide (NF_3O) F-1058
F3O1P1(g) Phosphoryl Fluoride (POF_3) F-1059
F3P1(g) Phosphorus Fluoride (PF_3) F-1060
F3P1S1(g) Thiophosphoryl Fluoride (PSF_3) F-1061
F3S1(g) Sulfur Fluoride (SF_3) F-1062
F3S1-(g) Sulfur Fluoride, Ion (SF_3^-) F-1063
F3S1+(g) Sulfur Fluoride, Ion (SF_3^+) F-1064
F3Si1(g) Trifluorosilyl (SiF_3) F-1065
F3Ti1(cr) Titanium Fluoride (TiF_3) F-1066
F3Ti1(g) Titanium Fluoride (TiF_3) F-1067
F3Zr1(cr) Zirconium Fluoride (ZrF_3) F-1068
F3Zr1(g) Zirconium Fluoride (ZrF_3) F-1069
F4H4(g) Hydrogen Fluoride (H_4F_4) F-1070
F4Mg2(g) Magnesium Fluoride ($(MgF_2)_2$) F-1071

F4Mo1(g) Molybdenum Fluoride (MoF_4) F-1072
F4Mo1O1(g) Molybdenum Fluoride Oxide (MoF_4O) F-1073
F4N2(g) Tetrafluorohydrazine (N_2F_4) F-1074
F4O1W1(cr) Tungsten Fluoride Oxide (WF_4O) F-1075
F4O1W1(cr,l) Tungsten Fluoride Oxide (WF_4O) F-1076
F4O1W1(g) Tungsten Fluoride Oxide (WF_4O) F-1077
F4O1W1(l) Tungsten Fluoride Oxide (WF_4O) F-1078
F4Pb1(g) Lead Fluoride (PbF_4) F-1079
F4S1(g) Sulfur Fluoride (SF_4) F-1080
F4S1-(g) Sulfur Fluoride, Ion (SF_4^-) F-1081
F4S1+(g) Sulfur Fluoride, Ion (SF_4^+) F-1082
F4Si1(g) Tetrafluorosilane (SiF_4) F-1083
F4Ti1(cr) Titanium Fluoride (TiF_4) F-1084
F4Ti1(g) Titanium Fluoride (TiF_4) F-1085
F4Zr1(cr) Zirconium Fluoride (ZrF_4) F-1086
F4Zr1(g) Zirconium Fluoride (ZrF_4) F-1087
F5H5(g) Hydrogen Fluoride (H_5F_5) F-1088
F5I1(g) Iodine Fluoride (IF_5) F-1089
F5Mo1(g) Molybdenum Fluoride (MoF_5) F-1090
F5P1(g) Phosphorus Fluoride (PF_5) F-1091
F5S1(g) Sulfur Fluoride (SF_5) F-1092
F5S1-(g) Sulfur Fluoride, Ion (SF_5^-) F-1093
F5S1+(g) Sulfur Fluoride, Ion (SF_5^+) F-1094
F6H6(g) Hydrogen Fluoride (H_6F_6) F-1095
F6Mo1(g) Molybdenum Fluoride (MoF_6) F-1096
F6Mo1(l) Molybdenum Fluoride (MoF_6) F-1097
F6S1(g) Sulfur Fluoride (SF_6) F-1098
F6S1-(g) Sulfur Fluoride, Ion (SF_6^-) F-1099
F6W1(g) Tungsten Fluoride (WF_6) F-1100
F6W1(l) Tungsten Fluoride (WF_6) F-1101
F7H7(g) Hydrogen Fluoride (H_7F_7) F-1102
F7I1(g) Iodine Fluoride (IF_7) F-1103
Fe0.877S1(cr) Iron Sulfide, Pyrrhotite ($Fe_{0.877}S$) F-1104
Fe0.947O1(cr) Iron Oxide, Wustite ($Fe_{0.947}O$) F-1105
Fe1(cr)_01 Iron, Alpha-Delta (Fe) F-1106
Fe1(cr)_02 Iron, Gamma (Fe) F-1107
Fe1(cr,l) Iron (Fe) F-1108
Fe1(g) Iron (Fe) F-1109
Fe1-(g) Iron, Ion (Fe^-) F-1110
Fe1(l) Iron (Fe) F-1111
Fe1+(g) Iron, Ion (Fe^+) F-1112
Fe1H2O2(cr) Iron Hydroxide ($Fe(OH)_2$) F-1113
Fe1H2O2(g) Iron Hydroxide ($Fe(OH)_2$) F-1114
Fe1H3O3(cr) Iron Hydroxide ($Fe(OH)_3$) F-1115
Fe1I2(cr) Iron Iodide (FeI_2) F-1116
Fe1I2(cr,l) Iron Iodide (FeI_2) F-1117

Fe1I2(g)	Iron Iodide (FeI_2)	F-1118
Fe1I2(l)	Iron Iodide (FeI_2)	F-1119
Fe1O1(cr)	Iron Oxide (FeO)	F-1120
Fe1O1(cr,l)	Iron Oxide (FeO)	F-1121
Fe1O1(g)	Iron Oxide (FeO)	F-1122
Fe1O1(l)	Iron Oxide (FeO)	F-1123
Fe1O4S1(cr)	Iron Sulfate ($FeSO_4$)	F-1124
Fe1S1(cr)	Iron Sulfide, Troilite (FeS)	F-1125
Fe1S1(cr,l)	Iron Sulfide, Troilite (FeS)	F-1126
Fe1S1(g)	Iron Sulfide (FeS)	F-1127
Fe1S1(l)	Iron Sulfide (FeS)	F-1128
Fe1S2(cr)_01	Iron Sulfide, Marcasite (FeS_2)	F-1129
Fe1S2(cr)_02	Iron Sulfide, Pyrite (FeS_2)	F-1130
Fe2I4(g)	Iron Iodide ($(FeI_2)_2$)	F-1131
Fe2O12S3(cr)	Iron Sulfate ($Fe_2(SO_4)_3$)	F-1132
Fe2O3(cr)	Iron Oxide, Hematite (Fe_2O_3)	F-1133
Fe3O4(cr)	Iron Oxide, Magnetite (Fe_3O_4)	F-1134
Ga1(cr)	Gallium (Ga)	F-1135
Ga1(cr,l)	Gallium (Ga)	F-1136
Ga1(g)	Gallium (Ga)	F-1137
Ga1-(g)	Gallium, Ion (Ga^-)	F-1138
Ga1(l)	Gallium (Ga)	F-1139
Ga1+(g)	Gallium, Ion (Ga^+)	F-1140
H1(g)	Hydrogen (H)	F-1141
H1-(g)	Hydrogen, Ion (H^-)	F-1142
H1+(g)	Hydrogen, Ion (H^+)	F-1143
H10O8S1(cr,l)	Sulfuric Acid, Tetrahydrate ($H_2SO_4 \cdot 4H_2O$)	F-1144
H15O10.5S1(cr,l)	Sulfuric Acid, Hemihexahydrate ($H_2SO_4 \cdot 6.5H_2O$)	F-1145
H1Hg1(g)	Mercury Hydride (HgH)	F-1146
H1I1(g)	Hydrogen Iodide (HI)	F-1147
H1I3Si1(g)	Triiodosilane ($SiHI_3$)	F-1148
H1K1(cr)	Potassium Hydride (KH)	F-1149
H1K1(g)	Potassium Hydride (KH)	F-1150
H1K1O1(cr)	Potassium Hydroxide (KOH)	F-1151
H1K1O1(cr,l)	Potassium Hydroxide (KOH)	F-1152
H1K1O1(g)	Potassium Hydroxide (KOH)	F-1153
H1K1O1(l)	Potassium Hydroxide (KOH)	F-1154
H1K1O1+(g)	Potassium Hydroxide, Ion (KOH^+)	F-1155
H1Li1(cr)	Lithium Hydride (LiH)	F-1156
H1Li1(cr,l)	Lithium Hydride (LiH)	F-1157
H1Li1(g)	Lithium Hydride (LiH)	F-1158
H1Li1(l)	Lithium Hydride (LiH)	F-1159
H1Li1O1(cr)	Lithium Hydroxide (LiOH)	F-1160
H1Li1O1(cr,l)	Lithium Hydroxide (LiOH)	F-1161
H1Li1O1(g)	Lithium Hydroxide (LiOH)	F-1162
H1Li1O1(l)	Lithium Hydroxide (LiOH)	F-1163

H1Li1O1+(g) Lithium Hydroxide, Ion ($LiOH^+$) F-1164
H1Mg1(g) Magnesium Hydride (MgH) F-1165
H1Mg1O1(g) Magnesium Hydroxide (MgOH) F-1166
H1Mg1O1+(g) Magnesium Hydroxide, Ion ($MgOH^+$) F-1167
H1N1(g) Imidogen (NH) F-1168
H1N1O1(g) Nitrosyl Hydride (HNO) F-1169
H1N1O2(g)_01 Nitrous Acid, Cis (HONO) F-1170
H1N1O2(g)_02 Nitrous Acid, Trans (HONO) F-1171
H1N1O3(g) Nitric Acid ($HONO_2$) F-1172
H1Na1(cr) Sodium Hydride (NaH) F-1173
H1Na1(g) Sodium Hydride (NaH) F-1174
H1Na1O1(cr) Sodium Hydroxide (NaOH) F-1175
H1Na1O1(cr,l) Sodium Hydroxide (NaOH) F-1176
H1Na1O1(g) Sodium Hydroxide (NaOH) F-1177
H1Na1O1(l) Sodium Hydroxide (NaOH) F-1178
H1Na1O1+(g) Sodium Hydroxide, Ion ($NaOH^+$) F-1179
H1O1(g) Hydroxyl (OH) F-1180
H1O1-(g) Hydroxyl, Ion (OH^-) F-1181
H1O1+(g) Hydroxyl, Ion (OH^+) F-1182
H1O1Sr1(g) Strontium Hydroxide (SrOH) F-1183
H1O1Sr1+(g) Strontium Hydroxide, Ion ($SrOH^+$) F-1184
H1O2(g) Hydroperoxo (HOO) F-1185
H1P1(g) Phosphinidene (PH) F-1186
H1Pb1(g) Lead Hydride (PbH) F-1187
H1S1(g) Mercapto (HS) F-1188
H1Si1(g) Silylidyne (SiH) F-1189
H1Si1+(g) Silylidyne, Ion (SiH^+) F-1190
H1Zr1(g) Zirconium Hydride (ZrH) F-1191
H2-(g) Hydrogen, Ion (H_2^-) F-1192
H2(ref) Hydrogen (H_2) F-1193
H2+(g) Hydrogen, Ion (H_2^+) F-1194
H2I2Si1(g) Diiodosilane (SiH_2I_2) F-1195
H2K2O2(g) Potassium Hydroxide ($(KOH)_2$) F-1196
H2Li2O2(g) Lithium Hydroxide ($(LiOH)_2$) F-1197
H2Mg1(cr) Magnesium Hydride (MgH_2) F-1198
H2Mg1O2(cr) Magnesium Hydroxide ($Mg(OH)_2$) F-1199
H2Mg1O2(g) Magnesium Hydroxide ($Mg(OH)_2$) F-1200
H2Mo1O4(g) Molybdic Acid ($O_2Mo(OH)_2$) F-1201
H2N1(g) Amidogen (NH_2) F-1202
H2N2(g) Diazene, Cis (HNNH) F-1203
H2Na2O2(g) Sodium Hydroxide ($(NaOH)_2$) F-1204
H2O1(fl)_01 Water, $_{500}$ Bar (H_2O) F-1205
H2O1(fl)_02 Water, $_{5000}$ Bar (H_2O) F-1206
H2O1(g) Water (H_2O) F-1207
H2O1(l) Water (H_2O) F-1208
H2O1(l,g)_$_{01}$ Water, $_1$ Bar (H2O) F-1209

H2O1(l,g)_0_2	Water, $_{10}$ Bar (H_2O)	F-1210
H2O1(l,g)_03	Water, $_{100}$ Bar (H_2O)	F-1211
H2O2(g)	Hydrogen Peroxide (HOOH)	F-1212
H2O2Sr1(cr)	Strontium Hydroxide ($Sr(OH)_2$)	F-1213
H2O2Sr1(cr,l)	Strontium Hydroxide ($Sr(OH)_2$)	F-1214
H2O2Sr1(g)	Strontium Hydroxide ($Sr(OH)_2$)	F-1215
H2O2Sr1(l)	Strontium Hydroxide ($Sr(OH)_2$)	F-1216
H2O4S1(cr,l)	Sulfuric Acid ($O_2S(OH)_2$)	F-1217
H2O4S1(g)	Sulfuric Acid ($O_2S(OH)_2$)	F-1218
H2O4W1(cr)	Tungstic Acid ($O_2W(OH)_2$)	F-1219
H2O4W1(g)	Tungstic Acid ($O_2W(OH)_2$)	F-1220
H2P1(g)	Phosphino (PH_2)	F-1221
H2S1(g)	Hydrogen Sulfide (H_2S)	F-1222
H2Ti1(cr)	Titanium Hydride (TiH_2)	F-1223
H3I1Si1(g)	Iodosilane (SiH_3I)	F-1224
H3N1(g)	Ammonia (NH_3)	F-1225
H3O1+(g)	Hydronium, Ion (H_3O^+)	F-1226
H3O4P1(cr)	Phosphoric Acid (H_3PO_4)	F-1227
H3O4P1(cr,l)	Phosphoric Acid (H_3PO_4)	F-1228
H3O4P1(l)	Phosphoric Acid (H_3PO_4)	F-1229
H3P1(g)	Phosphine (PH_3)	F-1230
H4I1N1(cr)	Ammonium Iodide (NH_4I)	F-1231
H4N2(g)	Hydrazine (N_2H_4)	F-1232
H4N2(l)	Hydrazine (N_2H_4)	F-1233
H4O5S1(cr,l)	Sulfuric Acid, Monohydrate ($H_2SO_4.H_2O$)	F-1234
H4Si1(g)	Silane (SiH_4)	F-1235
H6O6S1(cr,l)	Sulfuric Acid, Dihydrate ($H_2SO_{4.2}H_2O$)	F-1236
H8O7S1(cr,l)	Sulfuric Acid, Trihydrate ($H_2SO_{4.3}H_2O$)	F-1237
He1(ref)	Helium (He)	F-1238
He1+(g)	Helium, Ion (He^+)	F-1239
Hf1(cr)_01	Hafnium, Alpha (Hf)	F-1240
Hf1(cr)_02	Hafnium, Beta (Hf)	F-1241
Hf1(cr,l)	Hafnium (Hf)	F-1242
Hf1(g)	Hafnium (Hf)	F-1243
Hf1-(g)	Hafnium, Ion (Hf^-)	F-1244
Hf1(l)	Hafnium (Hf)	F-1245
Hf1+(g)	Hafnium, Ion (Hf^+)	F-1246
Hg1(cr,l)	Mercury (Hg)	F-1247
Hg1(g)	Mercury (Hg)	F-1248
Hg1+(g)	Mercury, Ion (Hg^+)	F-1249
Hg1I1(g)	Mercury Iodide (HgI)	F-1250
Hg1I2(cr)	Mercury Iodide (HgI_2)	F-1251
Hg1I2(cr,l)	Mercury Iodide (HgI_2)	F-1252
Hg1I2(g)	Mercury Iodide (HgI_2)	F-1253
Hg1I2(l)	Mercury Iodide (HgI_2)	F-1254
Hg1O1(cr)	Mercury Oxide (HgO)	F-1255

Hg1O1(g) Mercury Oxide (HgO) F-1256
Hg2I2(cr) Mercury Iodide (Hg_2I_2) F-1257
Hg2I2(cr,l) Mercury Iodide (Hg_2I_2) F-1258
Hg2I2(l) Mercury Iodide (Hg_2I_2) F-1259
I1(g) Iodine (I) F-1260
I1-(g) Iodine, Ion (I^-) F-1261
I1+(g) Iodine, Ion (I^+) F-1262
I1K1(cr) Potassium Iodide (KI) F-1263
I1K1(cr,l) Potassium Iodide (KI) F-1264
I1K1(g) Potassium Iodide (KI) F-1265
I1K1(l) Potassium Iodide (KI) F-1266
I1Li1(cr) Lithium Iodide (LiI) F-1267
I1Li1(cr,l) Lithium Iodide (LiI) F-1268
I1Li1(g) Lithium Iodide (LiI) F-1269
I1Li1(l) Lithium Iodide (LiI) F-1270
I1Mg1(g) Magnesium Iodide (MgI) F-1271
I1Mo1(g) Molybdenum Iodide (MoI) F-1272
I1N1O1(g) Nitrosyl Iodide (ONI) F-1273
I1Na1(cr) Sodium Iodide (NaI) F-1274
I1Na1(cr,l) Sodium Iodide (NaI) F-1275
I1Na1(l) Sodium Iodide (NaI) F-1276
I1Pb1(g) Lead Iodide (PbI) F-1277
I1Si1(g) Iodosilylidyne (SiI) F-1278
I1Sr1(g) Strontium Iodide (SrI) F-1279
I1Ti1(g) Titanium Iodide (TiI) F-1280
I1Zr1(g) Zirconium Iodide (ZrI) F-1281
I2(cr) Iodine (I_2) F-1282
I2(cr,l) Iodine (I_2) F-1283
I2(g) Iodine (I_2) F-1284
I2(l) Iodine (I_2) F-1285
I2K2(g) Potassium Iodide ($(KI)_2$) F-1286
I2Li2(g) Lithium Iodide ($(LiI)_2$) F-1287
I2Mg1(cr) Magnesium Iodide (MgI_2) F-1288
I2Mg1(cr,l) Magnesium Iodide (MgI_2) F-1289
I2Mg1(g) Magnesium Iodide (MgI_2) F-1290
I2Mg1(l) Magnesium Iodide (MgI_2) F-1291
I2Mo1(cr) Molybdenum Iodide, Alpha (MoI_2) F-1292
I2Mo1(g) Molybdenum Iodide (MoI_2) F-1293
I2Pb1(cr) Lead Iodide (PbI_2) F-1294
I2Pb1(cr,l) Lead Iodide (PbI_2) F-1295
I2Pb1(g) Lead Iodide (PbI_2) F-1296
I2Pb1(l) Lead Iodide (PbI_2) F-1297
I2Si1(g) Diiodosilylene (SiI_2) F-1298
I2Sr1(cr) Strontium Iodide (SrI_2) F-1299
I2Sr1(cr,l) Strontium Iodide (SrI_2) F-1300
I2Sr1(g) Strontium Iodide (SrI_2) F-1301

I2Sr1(l)	Strontium Iodide (SrI_2)	F-1302
I2Ti1(cr)	Titanium Iodide (TiI_2)	F-1303
I2Ti1(g)	Titanium Iodide (TiI_2)	F-1304
I2Zr1(cr)	Zirconium Iodide (ZrI_2)	F-1305
I2Zr1(cr,l)	Zirconium Iodide (ZrI_2)	F-1306
I2Zr1(g)	Zirconium Iodide (ZrI_2)	F-1307
I2Zr1(l)	Zirconium Iodide (ZrI_2)	F-1308
I3Mo1(cr)	Molybdenum Iodide (MoI_3)	F-1309
I3Mo1(g)	Molybdenum Iodide (MoI_3)	F-1310
I3Si1(g)	Triiodosilyl (SiI_3)	F-1311
I3Ti1(cr)	Titanium Iodide (TiI_3)	F-1312
I3Ti1(g)	Titanium Iodide (TiI_3)	F-1313
I3Zr1(cr)	Zirconium Iodide (ZrI_3)	F-1314
I3Zr1(g)	Zirconium Iodide (ZrI_3)	F-1315
I4Mo1(cr)	Molybdenum Iodide (MoI_4)	F-1316
I4Mo1(g)	Molybdenum Iodide (MoI_4)	F-1317
I4Pb1(g)	Lead Iodide (PbI_4)	F-1318
I4Si1(cr)	Tetraiodosilane (SiI_4)	F-1319
I4Si1(cr,l)	Tetraiodosilane (SiI_4)	F-1320
I4Si1(g)	Tetraiodosilane (SiI_4)	F-1321
I4Si1(l)	Tetraiodosilane (SiI_4)	F-1322
I4Ti1(cr)	Titanium Iodide (TiI_4)	F-1323
I4Ti1(cr,l)	Titanium Iodide (TiI_4)	F-1324
I4Ti1(g)	Titanium Iodide (TiI_4)	F-1325
I4Ti1(l)	Titanium Iodide (TiI_4)	F-1326
I4Zr1(cr)	Zirconium Iodide (ZrI_4)	F-1327
I4Zr1(g)	Zirconium Iodide (ZrI_4)	F-1328
K1(cr)	Potassium (K)	F-1329
K1(cr,l)	Potassium (K)	F-1330
K1(g)	Potassium (K)	F-1331
K1-(g)	Potassium, Ion (K^-)	F-1332
K1(l)	Potassium (K)	F-1333
K1+(g)	Potassium, Ion (K^+)	F-1334
K1O1(g)	Potassium Oxide (KO)	F-1335
K1O1-(g)	Potassium Oxide, Ion (KO^-)	F-1336
K1O2(cr)	Potassium Superoxide (KO_2)	F-1337
K2(g)	Potassium (K_2)	F-1338
K2O1(cr)	Potassium Oxide (K_2O)	F-1339
K2O2(cr)	Potassium Peroxide (K_2O_2)	F-1340
K2O3Si1(cr)	Potassium Silicate (K_2SiO_3)	F-1341
K2O3Si1(cr,l)	Potassium Silicate (K_2SiO_3)	F-1342
K2O3Si1(l)	Potassium Silicate (K_2SiO_3)	F-1343
K2O4S1(cr)_01	Potassium Sulfate, Alpha (K_2SO_4)	F-1344
K2O4S1(cr)_02	Potassium Sulfate, Beta (K_2SO_4)	F-1345
K2O4S1(cr,l)	Potassium Sulfate (K_2SO_4)	F-1346
K2O4S1(g)	Potassium Sulfate (K_2SO_4)	F-1347

K2O4S1(l)	Potassium Sulfate (K_2SO_4)	F-1348
K2S1(cr)	Potassium Sulfide (K_2S)	F-1349
K2S1(cr,l)	Potassium Sulfide (K_2S)	F-1350
K2S1(l)	Potassium Sulfide (K_2S)	F-1351
Kr1(ref)	Krypton (Kr)	F-1352
Kr1+(g)	Krypton, Ion (Kr^+)	F-1353
Li1(cr)	Lithium (Li)	F-1354
Li1(cr,l)	Lithium (Li)	F-1355
Li1(g)	Lithium (Li)	F-1356
Li1-(g)	Lithium, Ion (Li^-)	F-1357
Li1(l)	Lithium (Li)	F-1358
Li1+(g)	Lithium, Ion (Li^+)	F-1359
Li1N1(g)	Lithium Nitride (LiN)	F-1360
Li1N1O1(g)	Lithium Oxynitride (LiON)	F-1361
Li1Na1O1(g)	Lithium Sodium Oxide (LiONa)	F-1362
Li1O1(g)	Lithium Oxide (LiO)	F-1363
Li1O1-(g)	Lithium Oxide, Ion (LiO^-)	F-1364
Li2(g)	Lithium (Li_2)	F-1365
Li2O1(cr)	Lithium Oxide (Li_2O)	F-1366
Li2O1(cr,l)	Lithium Oxide (Li_2O)	F-1367
Li2O1(g)	Lithium Oxide (Li_2O)	F-1368
Li2O1(l)	Lithium Oxide (Li_2O)	F-1369
Li2O2(cr)	Lithium Peroxide (Li_2O_2)	F-1370
Li2O2(g)	Lithium Oxide ($(LiO)_2$)	F-1371
Li2O3Si1(cr)	Lithium Silicate (Li_2SiO_3)	F-1372
Li2O3Si1(cr,l)	Lithium Silicate (Li_2SiO_3)	F-1373
Li2O3Si1(l)	Lithium Silicate (Li_2SiO_3)	F-1374
Li2O3Ti1(cr)	Lithium Titanium Oxide (Li_2TiO_3)	F-1375
Li2O3Ti1(cr,l)	Lithium Titanium Oxide (Li_2TiO_3)	F-1376
Li2O3Ti1(l)	Lithium Titanium Oxide (Li_2TiO_3)	F-1377
Li2O4S1(cr)_01	Lithium Sulfate, Alpha (Li_2SO_4)	F-1378
Li2O4S1(cr)_02	Lithium Sulfate, Beta (Li_2SO_4)	F-1379
Li2O4S1(cr,l)	Lithium Sulfate (Li_2SO_4)	F-1380
Li2O4S1(g)	Lithium Sulfate (Li_2SO_4)	F-1381
Li2O4S1(l)	Lithium Sulfate (Li_2SO_4)	F-1382
Li2O5Si2(cr)	Lithium Silicate ($Li_2Si_2O_5$)	F-1383
Li2O5Si2(cr,l)	Lithium Silicate ($Li_2Si_2O_5$)	F-1384
Li2O5Si2(l)	Lithium Silicate ($Li_2Si_2O_5$)	F-1385
Li3N1(cr)	Lithium Nitride (Li_3N)	F-1386
Mg1(cr)	Magnesium (Mg)	F-1387
Mg1(cr,l)	Magnesium (Mg)	F-1388
Mg1(g)	Magnesium (Mg)	F-1389
Mg1(l)	Magnesium (Mg)	F-1390
Mg1+(g)	Magnesium, Ion (Mg^+)	F-1391
Mg1N1(g)	Magnesium Nitride (MgN)	F-1392
Mg1O1(cr)	Magnesium Oxide (MgO)	F-1393

Mg1O1(cr,l)	Magnesium Oxide (MgO)	F-1394
Mg1O1(g)	Magnesium Oxide (MgO)	F-1395
Mg1O1(l)	Magnesium Oxide (MgO)	F-1396
Mg1O3Si1(cr)	Magnesium Silicate ($MgSiO_3$)	F-1397
Mg1O3Si1(cr,l)	Magnesium Silicate ($MgSiO_3$)	F-1398
Mg1O3Si1(l)	Magnesium Silicate ($MgSiO_3$)	F-1399
Mg1O3Ti1(cr)	Magnesium Titanium Oxide ($MgTiO_3$)	F-1400
Mg1O3Ti1(cr,l)	Magnesium Titanium Oxide ($MgTiO_3$)	F-1401
Mg1O3Ti1(l)	Magnesium Titanium Oxide ($MgTiO_3$)	F-1402
Mg1O4S1(cr)	Magnesium Sulfate ($MgSO_4$)	F-1403
Mg1O4S1(cr,l)	Magnesium Sulfate ($MgSO_4$)	F-1404
Mg1O4S1(l)	Magnesium Sulfate ($MgSO_4$)	F-1405
Mg1O4W1(cr)	Magnesium Tungsten Oxide ($MgWO_4$)	F-1406
Mg1O5Ti2(cr)	Magnesium Titanium Oxide ($MgTi_2O_5$)	F-1407
Mg1O5Ti2(cr,l)	Magnesium Titanium Oxide ($MgTi_2O_5$)	F-1408
Mg1O5Ti2(l)	Magnesium Titanium Oxide ($MgTi_2O_5$)	F-1409
Mg1S1(cr)	Magnesium Sulfide (MgS)	F-1410
Mg1S1(g)	Magnesium Sulfide (MgS)	F-1411
Mg2(g)	Magnesium (Mg_2)	F-1412
Mg2O4Si1(cr)	Magnesium Silicate (Mg_2SiO_4)	F-1413
Mg2O4Si1(cr,l)	Magnesium Silicate (Mg_2SiO_4)	F-1414
Mg2O4Si1(l)	Magnesium Silicate (Mg_2SiO_4)	F-1415
Mg2O4Ti1(cr)	Magnesium Titanium Oxide (Mg_2TiO_4)	F-1416
Mg2O4Ti1(cr,l)	Magnesium Titanium Oxide (Mg_2TiO_4)	F-1417
Mg2O4Ti1(l)	Magnesium Titanium Oxide (Mg_2TiO_4)	F-1418
Mg2Si1(cr)	Magnesium Silicide (Mg_2Si)	F-1419
Mg2Si1(cr,l)	Magnesium Silicide (Mg_2Si)	F-1420
Mg2Si1(l)	Magnesium Silicide (Mg_2Si)	F-1421
Mg3N2(cr)	Magnesium Nitride (Mg_3N_2)	F-1422
Mg3O8P2(cr)	Magnesium Phosphate ($Mg_3P_2O_8$)	F-1423
Mg3O8P2(cr,l)	Magnesium Phosphate ($Mg_3P_2O_8$)	F-1424
Mg3O8P2(l)	Magnesium Phosphate ($Mg_3P_2O_8$)	F-1425
Mn1(cr)	Manganese (Mn)	F-1426
Mn1(cr,l)	Manganese (Mn)	F-1427
Mn1(g)	Manganese (Mn)	F-1428
Mn1(l)	Manganese (Mn)	F-1429
Mn1+(g)	Manganese, Ion (Mn^+)	F-1430
Mo1(cr)	Molybdenum (Mo)	F-1431
Mo1(cr,l)	Molybdenum (Mo)	F-1432
Mo1(g)	Molybdenum (Mo)	F-1433
Mo1-(g)	Molybdenum, Ion (Mo^-)	F-1434
Mo1(l)	Molybdenum (Mo)	F-1435
Mo1+(g)	Molybdenum, Ion (Mo^+)	F-1436
Mo1O1(g)	Molybdenum Oxide (MoO)	F-1437
Mo1O2(cr)	Molybdenum Oxide (MoO_2)	F-1438
Mo1O2(g)	Molybdenum Oxide (MoO_2)	F-1439

Mo1O2.750(cr)	Molybdenum Oxide ($MoO_{2.750}$)	F-1440
Mo1O2.875(cr)	Molybdenum Oxide ($MoO_{2.875}$)	F-1441
Mo1O2.889(cr)	Molybdenum Oxide ($MoO_{2.889}$)	F-1442
Mo1O3(cr)	Molybdenum Oxide (MoO_3)	F-1443
Mo1O3(cr,l)	Molybdenum Oxide (MoO_3)	F-1444
Mo1O3(g)	Molybdenum Oxide (MoO_3)	F-1445
Mo1O3(l)	Molybdenum Oxide (MoO_3)	F-1446
Mo1S2(cr)	Molybdenum Sulfide (MoS_2)	F-1447
Mo2S3(cr)	Molybdenum Sulfide (Mo_2S_3)	F-1448
Mo2S3(cr,l)	Molybdenum Sulfide (Mo_2S_3)	F-1449
Mo2S3(l)	Molybdenum Sulfide (Mo_2S_3)	F-1450
N0.465V1(cr)	Vanadium Nitride ($VN_{0.465}$)	F-1451
N1(g)	Nitrogen (N)	F-1452
N1-(g)	Nitrogen, Ion (N^-)	F-1453
N1+(g)	Nitrogen, Ion (N^+)	F-1454
N1O1(g)	Nitrogen Oxide (NO)	F-1455
N1O1+(g)	Nitrogen Oxide, Ion (NO^+)	F-1456
N1O2(g)	Nitrogen Oxide (NO_2)	F-1457
N1O2-(g)	Nitrogen Oxide, Ion (ONO^-)	F-1458
N1O3(g)	Nitrogen Oxide (NO_3)	F-1459
N1P1(g)	Phosphorus Nitride (PN)	F-1460
N1S1(g)	Nitrogen Sulfide (NS)	F-1461
N1Si1(g)	Silicon Nitride (SiN)	F-1462
N1Si2(g)	Silicon Nitride (Si_2N)	F-1463
N1Ti1(cr)	Titanium Nitride (TiN)	F-1464
N1Ti1(cr,l)	Titanium Nitride (TiN)	F-1465
N1Ti1(l)	Titanium Nitride (TiN)	F-1466
N1V1(cr)	Vanadium Nitride (VN)	F-1467
N1V1(g)	Vanadium Nitride (VN)	F-1468
N1Zr1(cr)	Zirconium Nitride (ZrN)	F-1469
N1Zr1(cr,l)	Zirconium Nitride (ZrN)	F-1470
N1Zr1(g)	Zirconium Nitride (ZrN)	F-1471
N1Zr1(l)	Zirconium Nitride (ZrN)	F-1472
N2-(g)	Nitrogen, Ion (N_2^-)	F-1473
N2(ref)	Nitrogen (N_2)	F-1474
N2+(g)	Nitrogen, Ion (N_2^+)	F-1475
N2O1(g)	Nitrogen Oxide (N_2O)	F-1476
N2O1+(g)	Nitrogen Oxide, Ion (N_2O^+)	F-1477
N2O3(g)	Nitrogen Oxide (N_2O_3)	F-1478
N2O4(cr)	Nitrogen Oxide (N_2O_4)	F-1479
N2O4(cr,l)	Nitrogen Oxide (N_2O_4)	F-1480
N2O4(g)	Nitrogen Oxide (N_2O_4)	F-1481
N2O4(l)	Nitrogen Oxide (N_2O_4)	F-1482
N2O5(g)	Nitrogen Oxide (N_2O_5)	F-1483
N3(g)	Azide (N_3)	F-1484
N4Si3(cr)	Silicon Nitride, Alpha (Si_3N_4)	F-1485

N5P3(cr)	Phosphorus Nitride (P_3N_5)	F-1486
Na1(cr)	Sodium (Na)	F-1487
Na1(cr,l)	Sodium (Na)	F-1488
Na1(g)	Sodium (Na)	F-1489
Na1-(g)	Sodium, Ion (Na^-)	F-1490
Na1(l)	Sodium (Na)	F-1491
Na1+(g)	Sodium, Ion (Na^+)	F-1492
Na1O1(g)	Sodium Oxide (NaO)	F-1493
Na1O1-(g)	Sodium Oxide, Ion (NaO^-)	F-1494
Na1O2(cr)	Sodium Superoxide (NaO_2)	F-1495
Na2(g)	Sodium (Na_2)	F-1496
Na2O1(cr)	Sodium Oxide (Na_2O)	F-1497
Na2O1(cr,l)	Sodium Oxide (Na_2O)	F-1498
Na2O1(l)	Sodium Oxide (Na_2O)	F-1499
Na2O2(cr)	Sodium Peroxide (Na_2O_2)	F-1500
Na2O3Si1(cr)	Sodium Silicate (Na_2SiO_3)	F-1501
Na2O3Si1(cr,l)	Sodium Silicate (Na_2SiO_3)	F-1502
Na2O3Si1(l)	Sodium Silicate (Na_2SiO_3)	F-1503
Na2O4S1(cr)_01	Sodium Sulfate, Delta (Na_2SO_4)	F-1504
Na2O4S1(cr)_02	Sodium Sulfate, I (Na_2SO_4)	F-1505
Na2O4S1(cr)_03	Sodium Sulfate, III (Na_2SO_4)	F-1506
Na2O4S1(cr)_04	Sodium Sulfate, IV (Na_2SO_4)	F-1507
Na2O4S1(cr)_05	Sodium Sulfate, V (Na_2SO_4)	F-1508
Na2O4S1(cr,l)	Sodium Sulfate (Na_2SO_4)	F-1509
Na2O4S1(g)	Sodium Sulfate (Na_2SO_4)	F-1510
Na2O4S1(l)	Sodium Sulfate (Na_2SO_4)	F-1511
Na2O4W1(cr)	Sodium Tungsten Oxide (Na_2WO_4)	F-1512
Na2O5Si2(cr)	Sodium Silicate ($Na_2Si_2O_5$)	F-1513
Na2O5Si2(cr,l)	Sodium Silicate ($Na_2Si_2O_5$)	F-1514
Na2O5Si2(l)	Sodium Silicate ($Na_2Si_2O_5$)	F-1515
Na2S1(cr)	Sodium Sulfide (Na_2S)	F-1516
Na2S1(cr,l)	Sodium Sulfide (Na_2S)	F-1517
Na2S1(l)	Sodium Sulfide (Na_2S)	F-1518
Na2S2(cr)	Sodium Sulfide, Beta (Na_2S_2)	F-1519
Na2S2(cr,l)	Sodium Sulfide (Na_2S_2)	F-1520
Na2S2(l)	Sodium Sulfide (Na_2S_2)	F-1521
Nb1(cr)	Niobium (Nb)	F-1522
Nb1(cr,l)	Niobium (Nb)	F-1523
Nb1(g)	Niobium (Nb)	F-1524
Nb1-(g)	Niobium, Ion (Nb^-)	F-1525
Nb1(l)	Niobium (Nb)	F-1526
Nb1+(g)	Niobium, Ion (Nb^+)	F-1527
Nb1O1(cr)	Niobium Oxide (NbO)	F-1528
Nb1O1(cr,l)	Niobium Oxide (NbO)	F-1529
Nb1O1(g)	Niobium Oxide (NbO)	F-1530
Nb1O1(l)	Niobium Oxide (NbO)	F-1531

Nb1O2(cr)	Niobium Oxide (NbO_2)	F-1532
Nb1O2(cr,l)	Niobium Oxide (NbO_2)	F-1533
Nb1O2(g)	Niobium Oxide (NbO_2)	F-1534
Nb1O2(l)	Niobium Oxide (NbO_2)	F-1535
Nb2O5(cr)	Niobium Oxide (Nb_2O_5)	F-1536
Nb2O5(cr,l)	Niobium Oxide (Nb_2O_5)	F-1537
Nb2O5(l)	Niobium Oxide (Nb_2O_5)	F-1538
Ne1(ref)	Neon (Ne)	F-1539
Ne1+(g)	Neon, Ion (Ne^+)	F-1540
Ni1(cr)	Nickel (Ni)	F-1541
Ni1(cr,l)	Nickel (Ni)	F-1542
Ni1(g)	Nickel (Ni)	F-1543
Ni1-(g)	Nickel, Ion (Ni^-)	F-1544
Ni1(l)	Nickel (Ni)	F-1545
Ni1+(g)	Nickel, Ion (Ni^+)	F-1546
Ni1S1(cr)	Nickel Sulfide (NiS)	F-1547
Ni1S1(cr,l)	Nickel Sulfide (NiS)	F-1548
Ni1S1(g)	Nickel Sulfide (NiS)	F-1549
Ni1S1(l)	Nickel Sulfide (NiS)	F-1550
Ni1S2(cr)	Nickel Sulfide (NiS_2)	F-1551
Ni1S2(cr,l)	Nickel Sulfide (NiS_2)	F-1552
Ni1S2(l)	Nickel Sulfide (NiS_2)	F-1553
Ni3S2(cr)	Nickel Sulfide (Ni_3S_2)	F-1554
Ni3S2(cr,l)	Nickel Sulfide (Ni_3S_2)	F-1555
Ni3S2(l)	Nickel Sulfide (Ni_3S_2)	F-1556
Ni3S4(cr)	Nickel Sulfide (Ni_3S_4)	F-1557
O1(g)	Oxygen (O)	F-1558
O1-(g)	Oxygen, Ion (O^-)	F-1559
O1+(g)	Oxygen, Ion (O^+)	F-1560
O10P4(cr)	Phosphorus Oxide ($(P_2O_5)_2$)	F-1561
O10P4(g)	Phosphorus Oxide ($(P_2O_5)_2$)	F-1562
O12W4(g)	Tungsten Oxide ($(WO_3)_4$)	F-1563
O1P1(g)	Phosphorus Oxide (PO)	F-1564
O1Pb1(cr)_01	Lead Oxide, Red (PbO)	F-1565
O1Pb1(cr)_02	Lead Oxide, Yellow (PbO)	F-1566
O1Pb1(cr,l)	Lead Oxide (PbO)	F-1567
O1Pb1(g)	Lead Oxide (PbO)	F-1568
O1Pb1(l)	Lead Oxide (PbO)	F-1569
O1S1(g)	Sulfur Oxide (SO)	F-1570
O1S2(g)	Sulfur Oxide (SSO)	F-1571
O1Si1(g)	Silicon Oxide (SiO)	F-1572
O1Sr1(cr)	Strontium Oxide (SrO)	F-1573
O1Sr1(cr,l)	Strontium Oxide (SrO)	F-1574
O1Sr1(g)	Strontium Oxide (SrO)	F-1575
O1Sr1(l)	Strontium Oxide (SrO)	F-1576
O1Ta1(g)	Tantalum Oxide (TaO)	F-1577

O1Ti1(cr)_01	Titanium Oxide, Alpha (TiO)	F-1578
O1Ti1(cr)_02	Titanium Oxide, Beta (TiO)	F-1579
O1Ti1(cr,l)	Titanium Oxide (TiO)	F-1580
O1Ti1(g)	Titanium Oxide (TiO)	F-1581
O1Ti1(l)	Titanium Oxide (TiO)	F-1582
O1V1(cr)	Vanadium Oxide (VO)	F-1583
O1V1(cr,l)	Vanadium Oxide (VO)	F-1584
O1V1(g)	Vanadium Oxide (VO)	F-1585
O1V1(l)	Vanadium Oxide (VO)	F-1586
O1W1(g)	Tungsten Oxide (WO)	F-1587
O1Zr1(g)	Zirconium Oxide (ZrO)	F-1588
O2-(g)	Oxygen, Ion (O_2^-)	F-1589
O2(ref)	Oxygen (O_2)	F-1590
O2.72W1(cr)	Tungsten Oxide ($WO_{2.72}$)	F-1591
O2.90W1(cr)	Tungsten Oxide ($WO_{2.90}$)	F-1592
O2.96W1(cr)	Tungsten Oxide ($WO_{2.96}$)	F-1593
O2+(g)	Oxygen, Ion (O_2^+)	F-1594
O2P1(g)	Phosphorus Oxide (PO_2)	F-1595
O2Pb1(cr)	Lead Oxide (PbO_2)	F-1596
O2S1(g)	Sulfur Dioxide (SO_2)	F-1597
O2Si1(cr)_01	Silicon Oxide, Cristobalite, High (SiO_2)	F-1598
O2Si1(cr)_02	Silicon Oxide, Cristobalite, Low (SiO_2)	F-1599
O2Si1(cr)_03	Silicon Oxide, Quartz (SiO_2)	F-1600
O2Si1(cr,l)	Silicon Oxide (SiO_2)	F-1601
O2Si1(g)	Silicon Oxide (SiO_2)	F-1602
O2Si1(l)	Silicon Oxide (SiO_2)	F-1603
O2Ta1(g)	Tantalum Oxide (TaO_2)	F-1604
O2Ti1(cr)_01	Titanium Oxide, Anatase (TiO_2)	F-1605
O2Ti1(cr)_02	Titanium Oxide, Rutile (TiO_2)	F-1606
O2Ti1(cr,l)	Titanium Oxide (TiO_2)	F-1607
O2Ti1(g)	Titanium Oxide (TiO_2)	F-1608
O2Ti1(l)	Titanium Oxide (TiO_2)	F-1609
O2V1(g)	Vanadium Oxide (VO_2)	F-1610
O2W1(cr)	Tungsten Oxide (WO_2)	F-1611
O2W1(g)	Tungsten Oxide (WO_2)	F-1612
O2Zr1(cr)	Zirconium Oxide (ZrO_2)	F-1613
O2Zr1(cr,l)	Zirconium Oxide (ZrO_2)	F-1614
O2Zr1(g)	Zirconium Oxide (ZrO_2)	F-1615
O2Zr1(l)	Zirconium Oxide (ZrO_2)	F-1616
O3(g)	Ozone (O_3)	F-1617
O3Pb1Si1(cr)	Lead Silicate ($PbSiO_3$)	F-1618
O3S1(g)	Sulfur Trioxide (SO_3)	F-1619
O3Ti2(cr)	Titanium Oxide (Ti_2O_3)	F-1620
O3Ti2(cr,l)	Titanium Oxide (Ti_2O_3)	F-1621
O3Ti2(l)	Titanium Oxide (Ti_2O_3)	F-1622
O3V2(cr)	Vanadium Oxide (V_2O_3)	F-1623

O3V2(cr,l)	Vanadium Oxide (V_2O_3)	F-1624
O3V2(l)	Vanadium Oxide (V_2O_3)	F-1625
O3W1(cr)	Tungsten Oxide (WO_3)	F-1626
O3W1(cr,l)	Tungsten Oxide (WO_3)	F-1627
O3W1(g)	Tungsten Oxide (WO_3)	F-1628
O3W1(l)	Tungsten Oxide (WO_3)	F-1629
O4Pb2Si1(cr)	Lead Silicate (Pb_2SiO_4)	F-1630
O4Pb3(cr)	Lead Oxide (Pb_3O_4)	F-1631
O4S1Zn1(cr)	Zinc Sulfate ($ZnSO_4$)	F-1632
O4Si1Zr1(cr)	Zirconium Silicate ($ZrSiO_4$)	F-1633
O4V2(cr)	Vanadium Oxide (V_2O_4)	F-1634
O4V2(cr,l)	Vanadium Oxide (V_2O_4)	F-1635
O4V2(l)	Vanadium Oxide (V_2O_4)	F-1636
O5Ta2(cr)	Tantalum Oxide (Ta_2O_5)	F-1637
O5Ta2(cr,l)	Tantalum Oxide (Ta_2O_5)	F-1638
O5Ta2(l)	Tantalum Oxide (Ta_2O_5)	F-1639
O5Ti3(cr)_01	Titanium Oxide, Alpha (Ti_3O_5)	F-1640
O5Ti3(cr)_02	Titanium Oxide, Beta (Ti_3O_5)	F-1641
O5Ti3(cr,l)	Titanium Oxide (Ti_3O_5)	F-1642
O5Ti3(l)	Titanium Oxide (Ti_3O_5)	F-1643
O5V2(cr)	Vanadium Oxide (V_2O_5)	F-1644
O5V2(cr,l)	Vanadium Oxide (V_2O_5)	F-1645
O5V2(l)	Vanadium Oxide (V_2O_5)	F-1646
O6P4(g)	Phosphorus Oxide ($(P_2O_3)_2$)	F-1647
O6W2(g)	Tungsten Oxide ($(WO_3)_2$)	F-1648
O7Ti4(cr)	Titanium Oxide (Ti_4O_7)	F-1649
O7Ti4(cr,l)	Titanium Oxide (Ti_4O_7)	F-1650
O7Ti4(l)	Titanium Oxide (Ti_4O_7)	F-1651
O8W3(g)	Tungsten Oxide (W_3O_8)	F-1652
O9W3(g)	Tungsten Oxide ($(WO_3)_3$)	F-1653
P1(cr)_01	Phosphorus, White (P)	F-1654
P1(cr)_02	Phosphorus, Red, IV (P)	F-1655
P1(cr)_03	Phosphorus, Red, V (P)	F-1656
P1(cr)_04	Phosphorus, Black (P)	F-1657
P1(cr,l)	Phosphorus (P)	F-1658
P1(g)	Phosphorus (P)	F-1659
P1-(g)	Phosphorus, Ion (P^-)	F-1660
P1(l)	Phosphorus (P)	F-1661
P1+(g)	Phosphorus, Ion (P^+)	F-1662
P1S1(g)	Phosphorus Sulfide (PS)	F-1663
P2(g)	Phosphorus (P_2)	F-1664
P4(g)	Phosphorus (P_4)	F-1665
P4S3(cr)	Phosphorus Sulfide (P_4S_3)	F-1666
P4S3(cr,l)	Phosphorus Sulfide (P_4S_3)	F-1667
P4S3(g)	Phosphorus Sulfide (P_4S_3)	F-1668
P4S3(l)	Phosphorus Sulfide (P_4S_3)	F-1669

Pb1(cr)	Lead (Pb)	F-1670
Pb1(cr,l)	Lead (Pb)	F-1671
Pb1(g)	Lead (Pb)	F-1672
Pb1-(g)	Lead, Ion (Pb^-)	F-1673
Pb1(l)	Lead (Pb)	F-1674
Pb1+(g)	Lead, Ion (Pb^+)	F-1675
Pb1S1(cr)	Lead Sulfide (PbS)	F-1676
Pb1S1(cr,l)	Lead Sulfide (PbS)	F-1677
Pb1S1(g)	Lead Sulfide (PbS)	F-1678
Pb1S1(l)	Lead Sulfide (PbS)	F-1679
Pb2(g)	Lead (Pb_2)	F-1680
Rb1(cr)	Rubidium (Rb)	F-1681
Rb1(cr,l)	Rubidium (Rb)	F-1682
Rb1(g)	Rubidium (Rb)	F-1683
Rb1-(g)	Rubidium, Ion (Rb^-)	F-1684
Rb1(l)	Rubidium (Rb)	F-1685
Rb1+(g)	Rubidium, Ion (Rb^+)	F-1686
Rb2(g)	Rubidium (Rb_2)	F-1687
Rn1(ref)	Radon (Rn)	F-1688
Rn1+(g)	Radon, Ion (Rn^+)	F-1689
S1(cr)_01	Sulfur, Orthorhombic (S)	F-1690
S1(cr)_02	Sulfur, Monoclinic (S)	F-1691
S1(cr,l)	Sulfur (S)	F-1692
S1(g)	Sulfur (S)	F-1693
S1-(g)	Sulfur, Ion (S^-)	F-1694
S1(l)	Sulfur (S)	F-1695
S1+(g)	Sulfur, Ion (S^+)	F-1696
S1Si1(g)	Silicon Sulfide (SiS)	F-1697
S1Sr1(cr)	Strontium Sulfide (SrS)	F-1698
S1Sr1(g)	Strontium Sulfide (SrS)	F-1699
S2(g)	Sulfur (S_2)	F-1700
S2Si1(cr)	Silicon Sulfide (SiS_2)	F-1701
S2Si1(cr,l)	Silicon Sulfide (SiS_2)	F-1702
S2Si1(l)	Silicon Sulfide (SiS_2)	F-1703
S3(g)	Sulfur (S_3)	F-1704
S4(g)	Sulfur (S_4)	F-1705
S5(g)	Sulfur (S_5)	F-1706
S6(g)	Sulfur (S_6)	F-1707
S7(g)	Sulfur (S_7)	F-1708
S8(g)	Sulfur (S_8)	F-1709
Si1(cr)	Silicon (Si)	F-1710
Si1(cr,l)	Silicon (Si)	F-1711
Si1(g)	Silicon (Si)	F-1712
Si1-(g)	Silicon, Ion (Si^-)	F-1713
Si1(l)	Silicon (Si)	F-1714
Si1+(g)	Silicon, Ion (Si^+)	F-1715

Si2(g)	Silicon (Si_2)	F-1716
Si3(g)	Silicon (Si_3)	F-1717
Sr1(cr)_01	Strontium, Alpha (Sr)	F-1718
Sr1(cr)_02	Strontium, Beta (Sr)	F-1719
Sr1(cr,l)	Strontium (Sr)	F-1720
Sr1(g)	Strontium (Sr)	F-1721
Sr1(l)	Strontium (Sr)	F-1722
Sr1+(g)	Strontium, Ion (Sr^+)	F-1723
Ta1(cr)	Tantalum (Ta)	F-1724
Ta1(cr,l)	Tantalum (Ta)	F-1725
Ta1(g)	Tantalum (Ta)	F-1726
Ta1-(g)	Tantalum, Ion (Ta^-)	F-1727
Ta1(l)	Tantalum (Ta)	F-1728
Ta1+(g)	Tantalum, Ion (Ta^+)	F-1729
Ti1(cr)_01	Titanium, Alpha (Ti)	F-1730
Ti1(cr)_02	Titanium, Beta (Ti)	F-1731
Ti1(cr,l)	Titanium (Ti)	F-1732
Ti1(g)	Titanium (Ti)	F-1733
Ti1-(g)	Titanium, Ion (Ti^-)	F-1734
Ti1(l)	Titanium (Ti)	F-1735
Ti1+(g)	Titanium, Ion (Ti^+)	F-1736
V1(cr)	Vanadium (V)	F-1737
V1(cr,l)	Vanadium (V)	F-1738
V1(g)	Vanadium (V)	F-1739
V1-(g)	Vanadium, Ion (V^-)	F-1740
V1(l)	Vanadium (V)	F-1741
V1+(g)	Vanadium, Ion (V^+)	F-1742
W1(cr)	Tungsten (W)	F-1743
W1(cr,l)	Tungsten (W)	F-1744
W1(g)	Tungsten (W)	F-1745
W1-(g)	Tungsten, Ion (W^-)	F-1746
W1(l)	Tungsten (W)	F-1747
W1+(g)	Tungsten, Ion (W^+)	F-1748
Xe1(ref)	Xenon (Xe)	F-1749
Xe1+(g)	Xenon, Ion (Xe^+)	F-1750
Zn1(cr)	Zinc (Zn)	F-1751
Zn1(cr,l)	Zinc (Zn)	F-1752
Zn1(g)	Zinc (Zn)	F-1753
Zn1-(g)	Zinc, Ion (Zn^-)	F-1754
Zn1(l)	Zinc (Zn)	F-1755
Zn1+(g)	Zinc, Ion (Zn^+)	F-1756
Zr1(cr)_01	Zirconium, Alpha (Zr)	F-1757
Zr1(cr)_02	Zirconium, Beta (Zr)	F-1758
Zr1(cr,l)	Zirconium (Zr)	F-1759
Zr1(g)	Zirconium (Zr)	F-1760
Zr1-(g)	Zirconium, Ion (Zr^-)	F-1761

Zr1(l) Zirconium (Zr) F-1762
Zr1+(g) Zirconium, Ion (Zr^+) F-1763

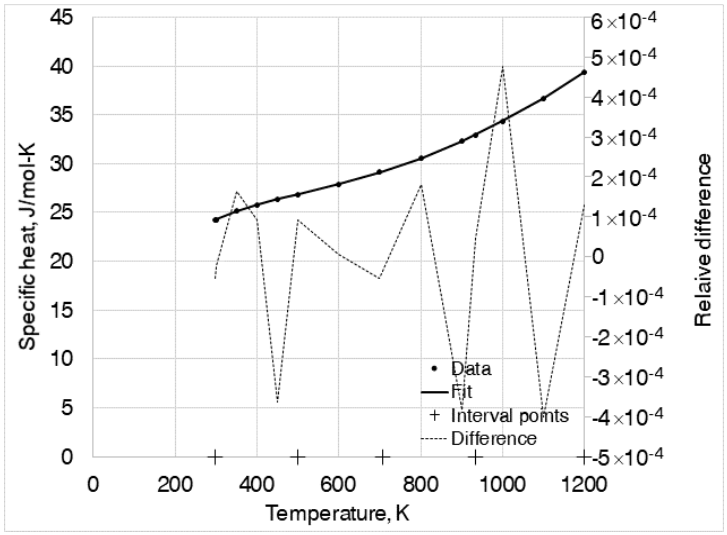
Appendix F
Plots of JANAF Data Fits

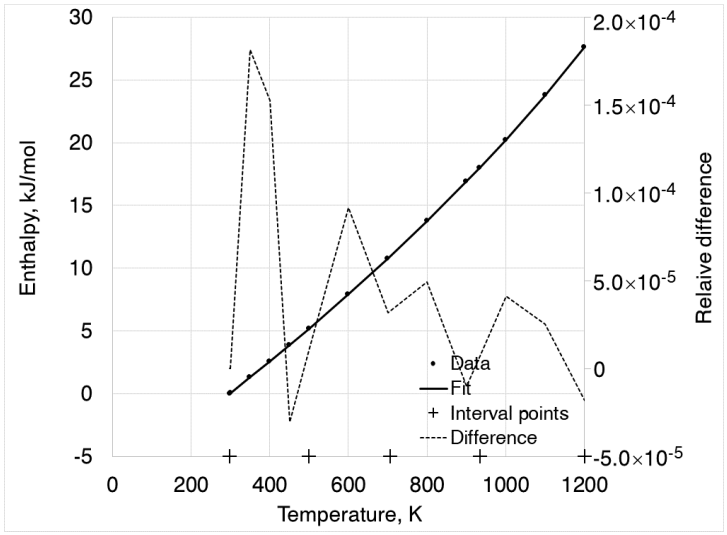
This appendix contains plots of the data fits compared to the tabular JANAF data for each of the species and phases listed in appendix B. Plots of specific heat, enthalpy, entropy, and reduced thermodynamic potential $-(G^0 - H^0(T_{ref}))/T$ are given as a function of temperature for a reference state of 0.1 MPa. The relative difference between the published JANAF data values and fits for each species is also shown.

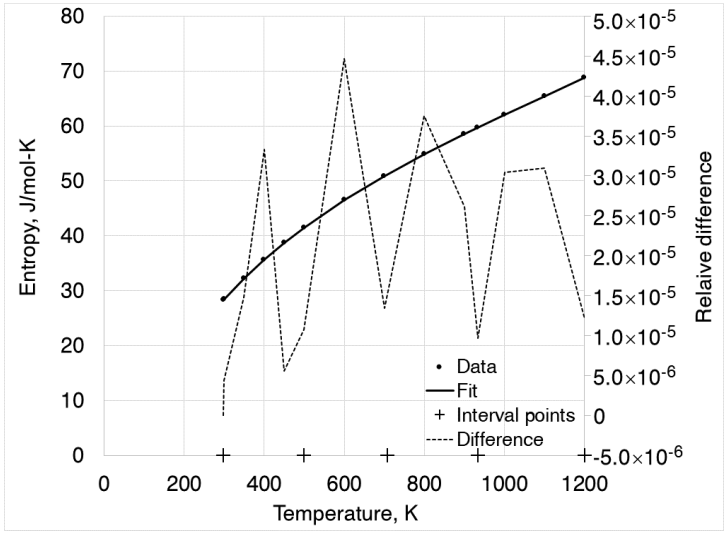
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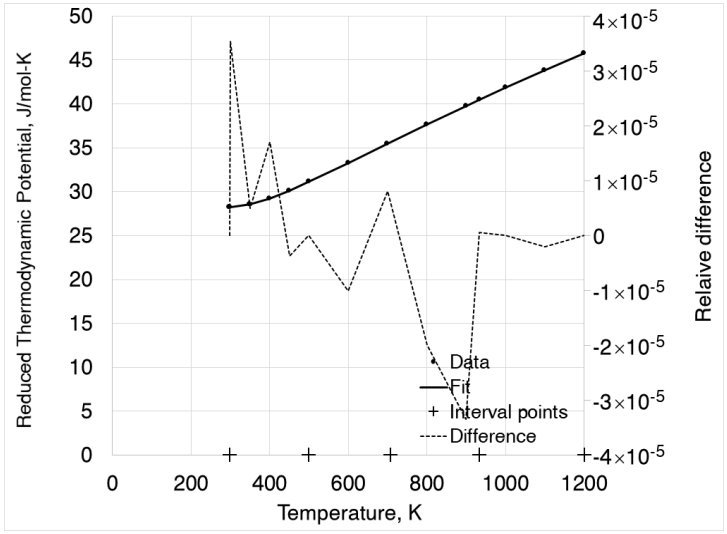
Aluminum (Al) – Al1(cr)

MW = 26.98154 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 28.275 J/mol-K – p^o = 0.1 MPa



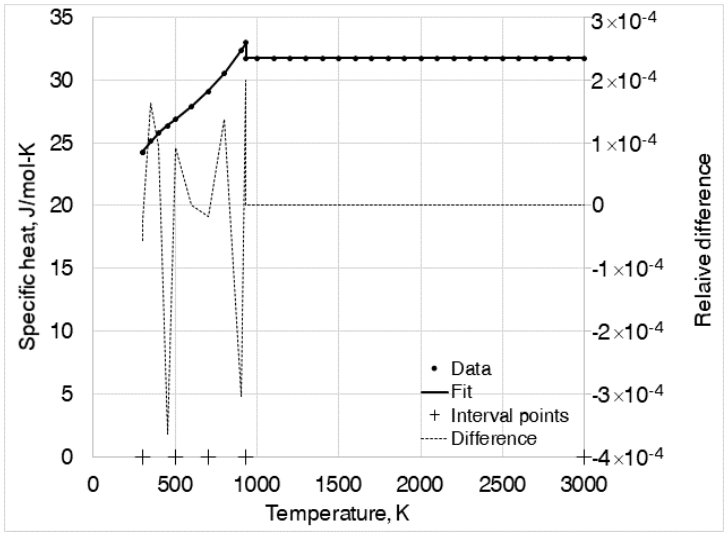


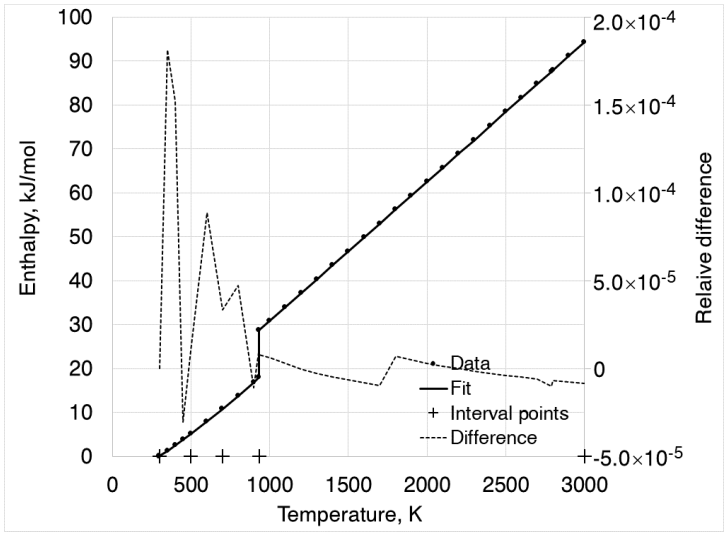


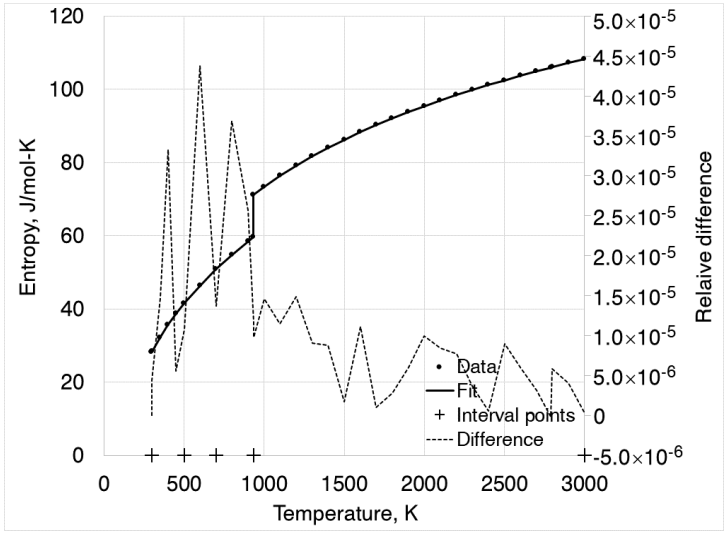


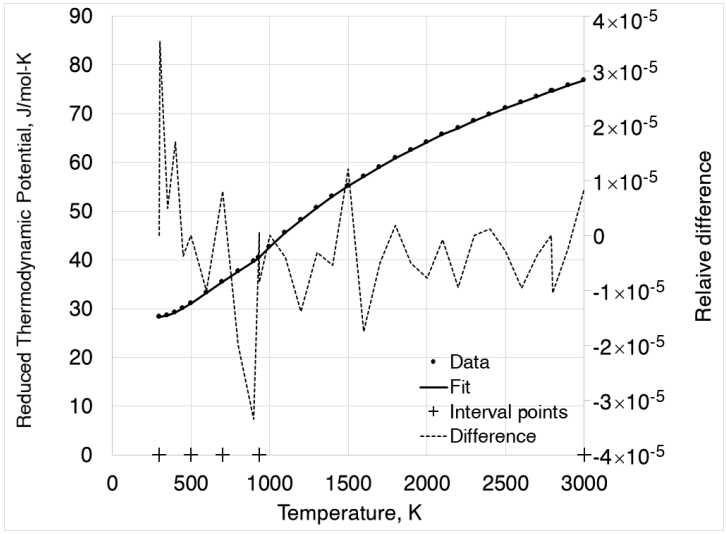
Aluminum (Al) – Al1(cr,l)

MW = 26.98154 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 28.275 J/mol-K – p^o = 0.1 MPa



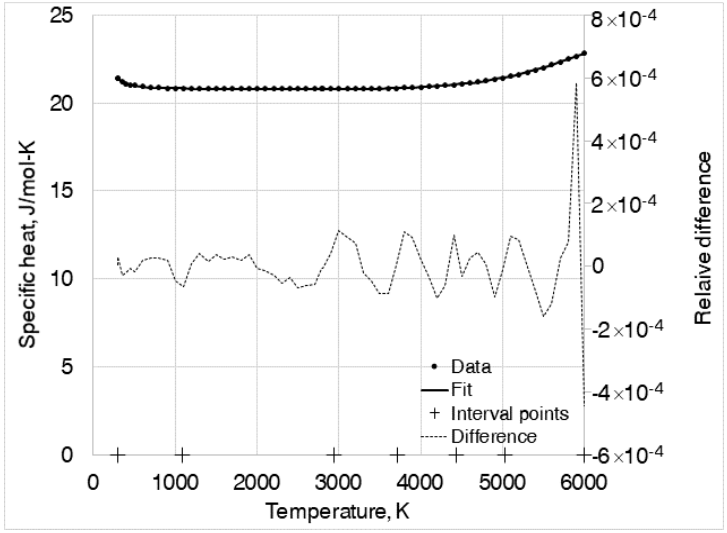


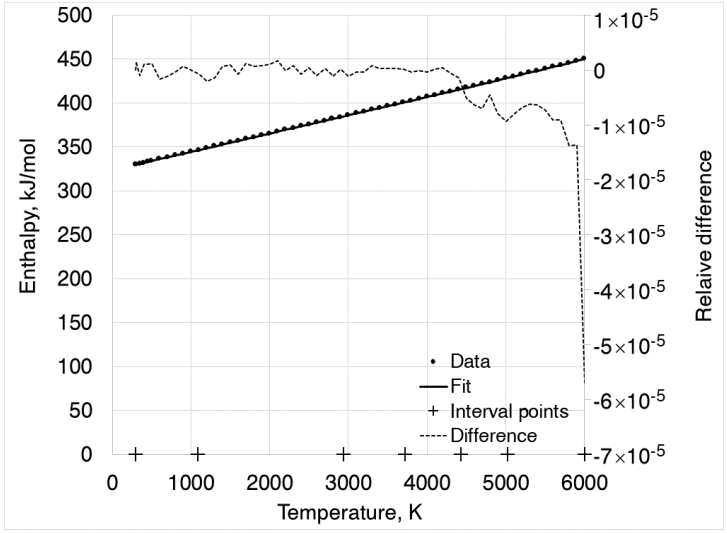


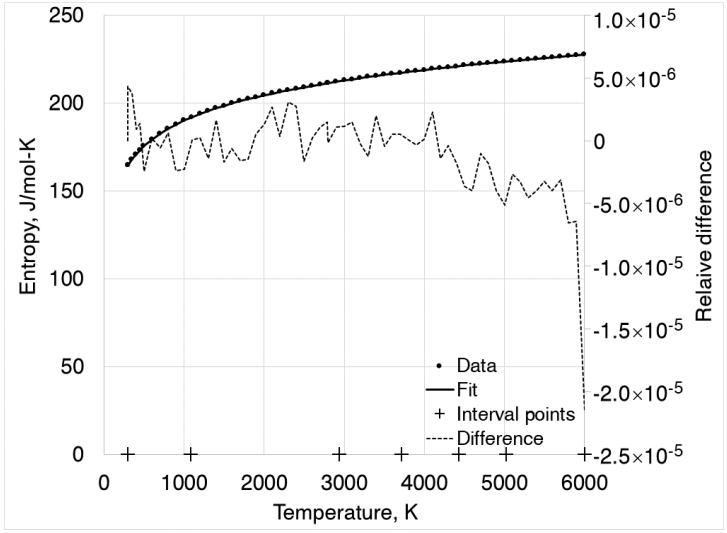


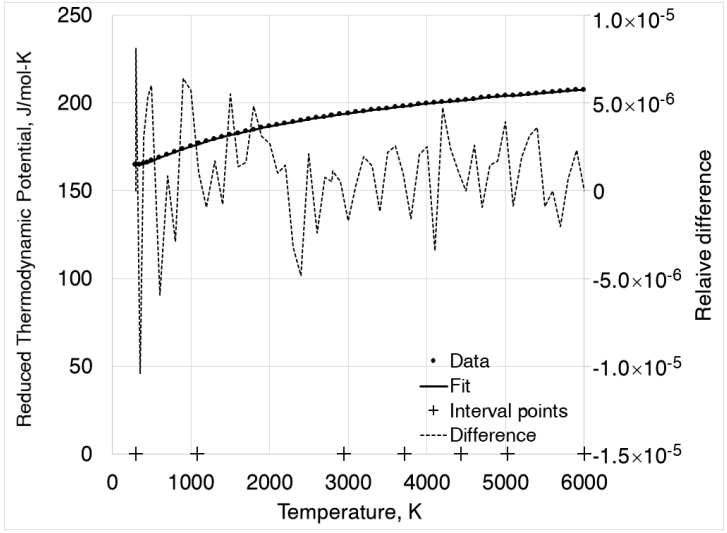
Aluminum (Al) – Al1(g)

MW = 26.98154 g/mol – $\Delta_f H^o$(298.15 K) = 329.699 kJ/mol – S^o(298.15 K) = 164.553 J/mol-K – p^o = 0.1 MPa



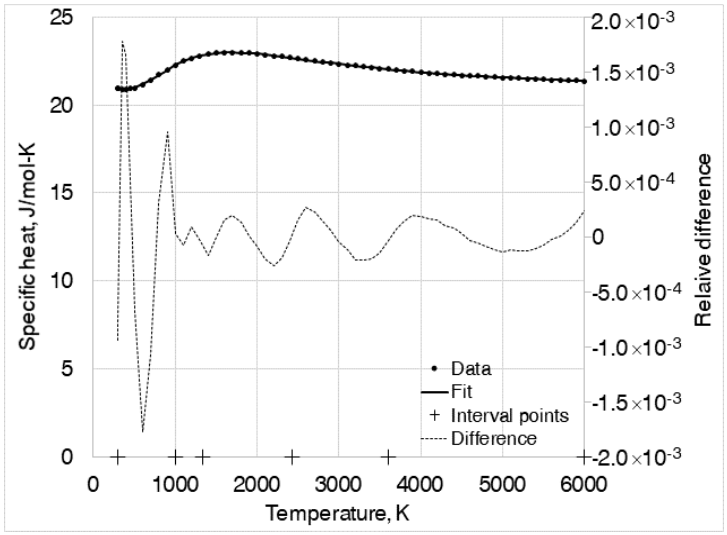


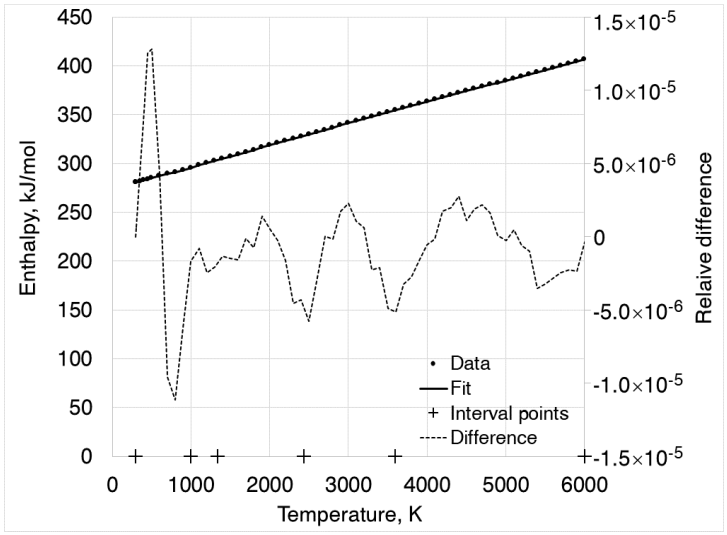


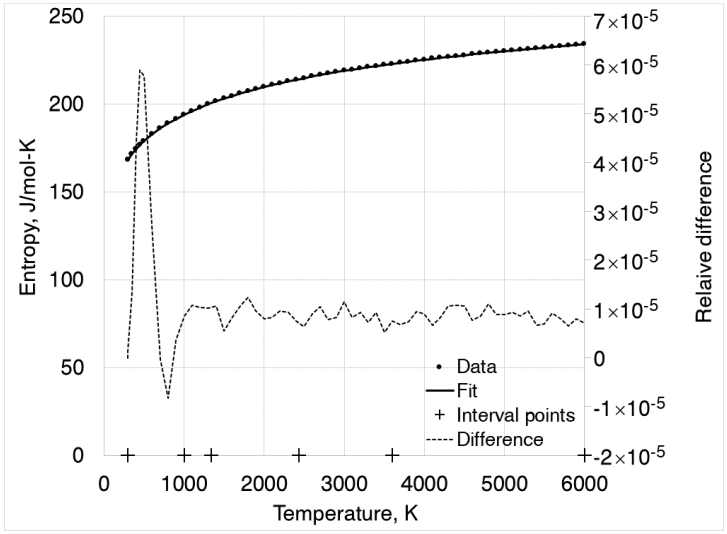


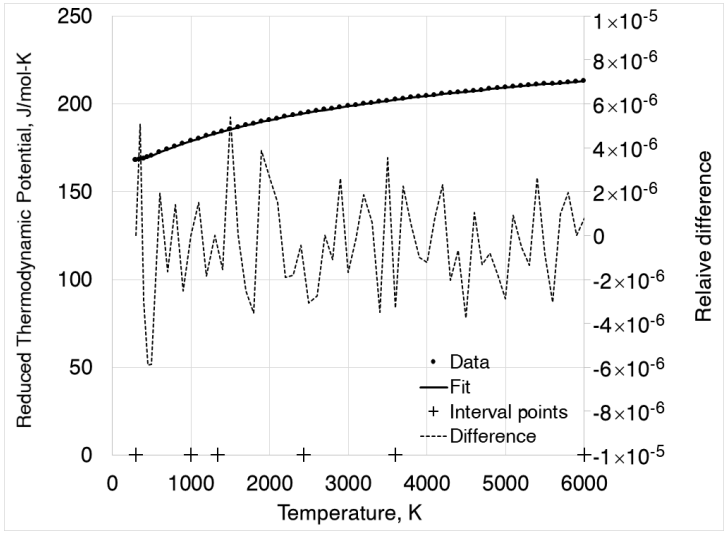
Aluminum, Ion (Al^-) – Al1-(g)

MW = 26.98209 g/mol – $\Delta_f H^o$(298.15 K) = 280.79 kJ/mol – S^o(298.15 K) = 168.135 J/mol-K – p^o = 0.1 MPa



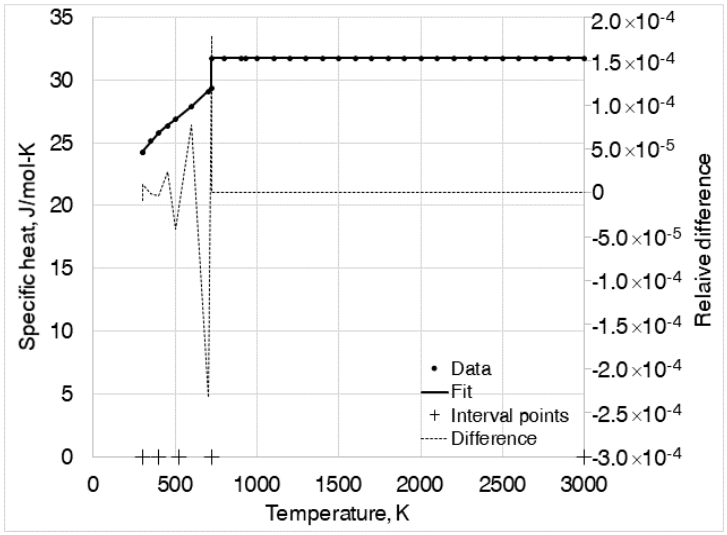


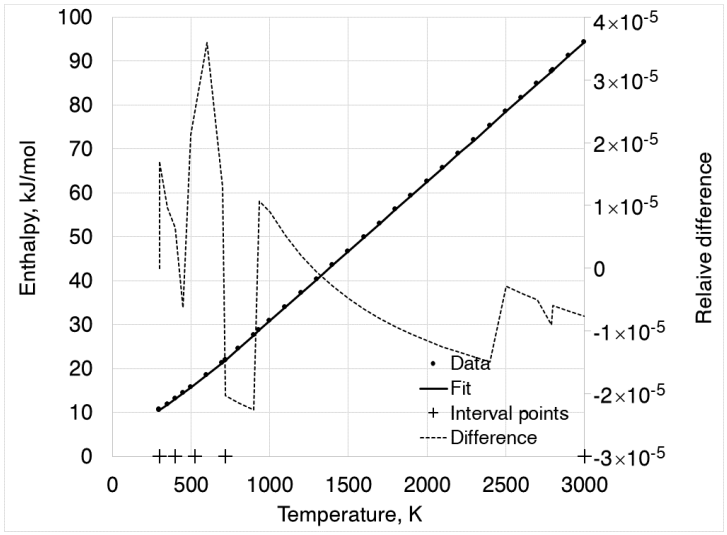


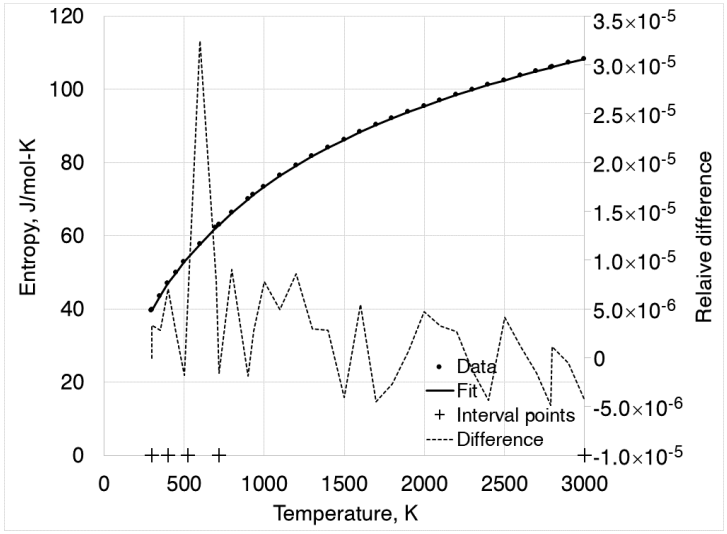


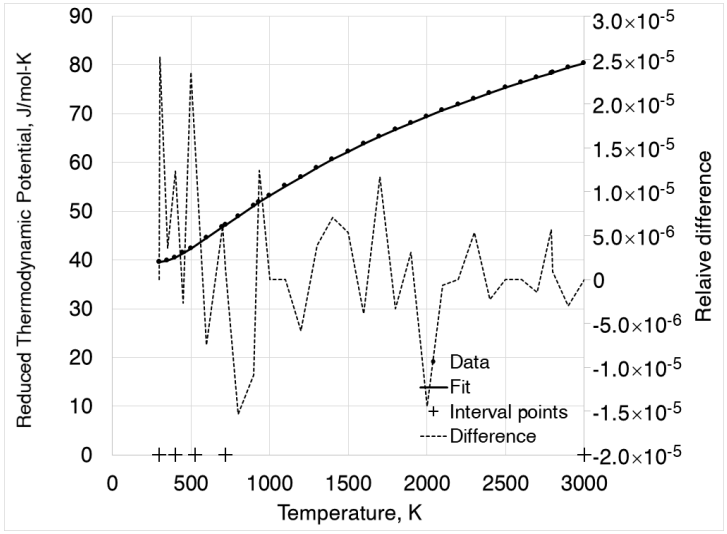
Aluminum (Al) – Al1(l)

MW = 26.98154 g/mol – $\Delta_f H^o$(298.15 K) = 10.562 kJ/mol – S^o(298.15 K) = 39.549 J/mol-K – p^o = 0.1 MPa



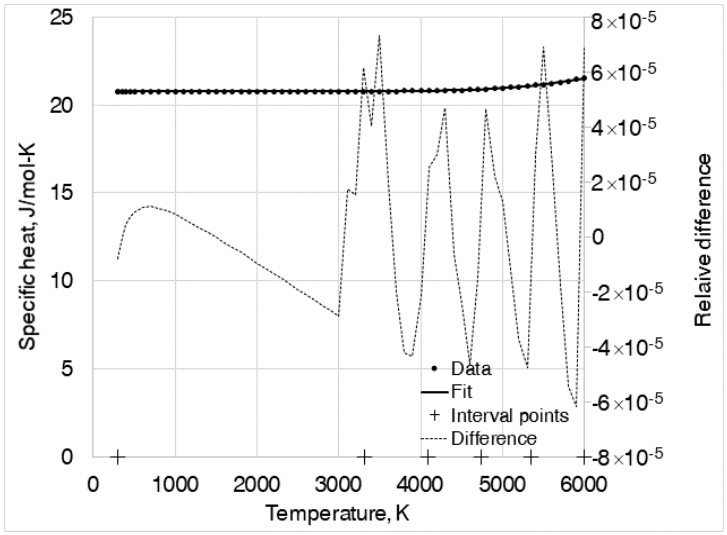


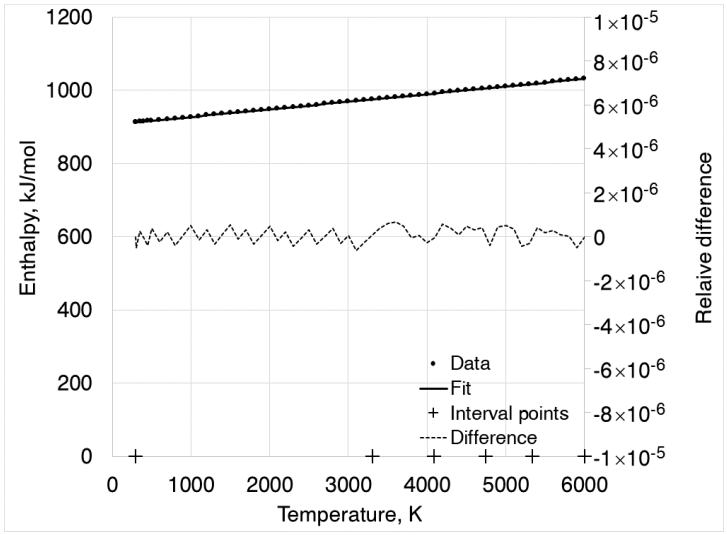


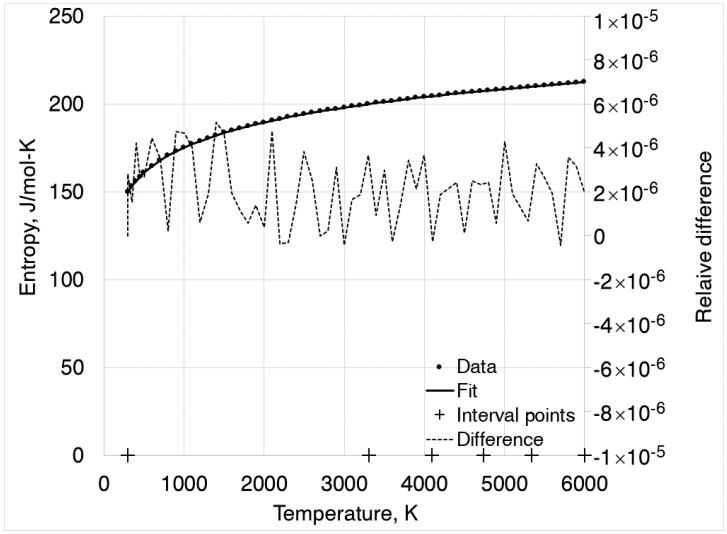


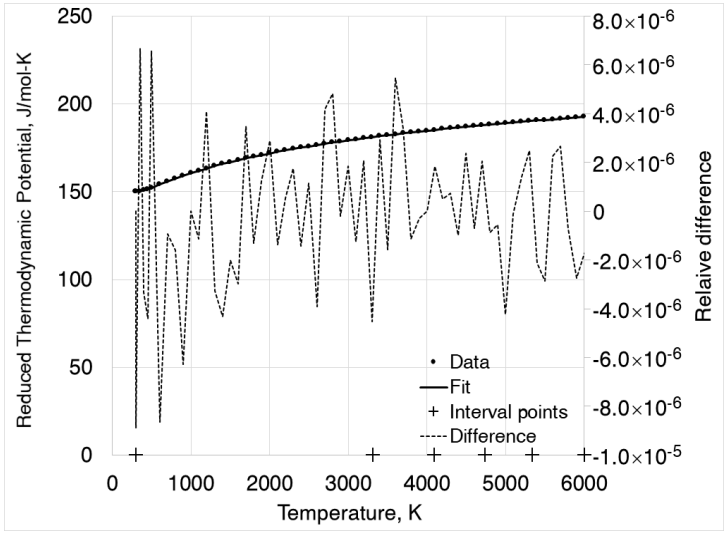
Aluminum, Ion (Al^+) – Al1+(g)

MW = 26.98099 g/mol – $\Delta_f H^o$(298.15 K) = 912.713 kJ/mol – S^o(298.15 K) = 149.95 J/mol-K – p^o = 0.1 MPa



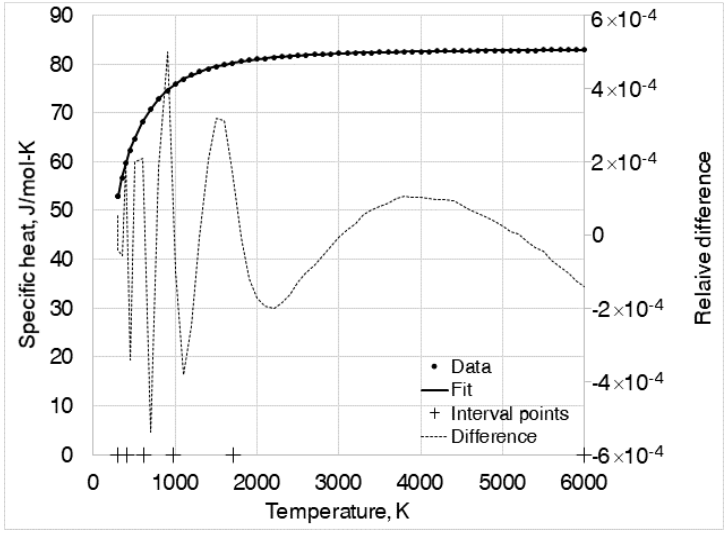


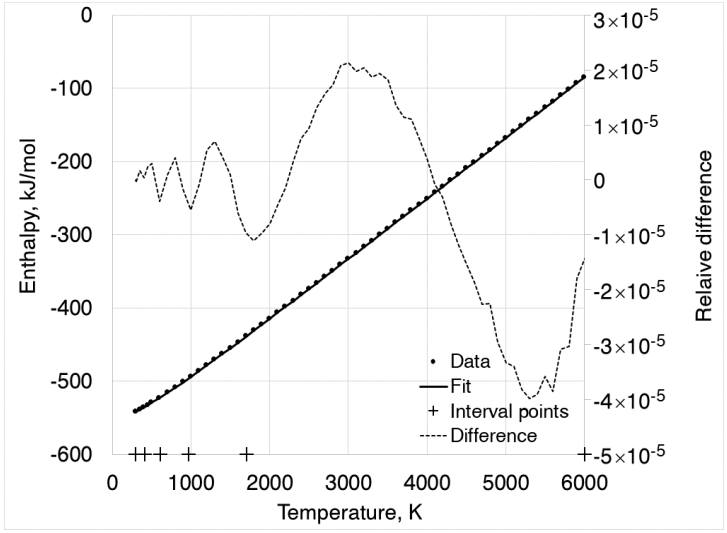


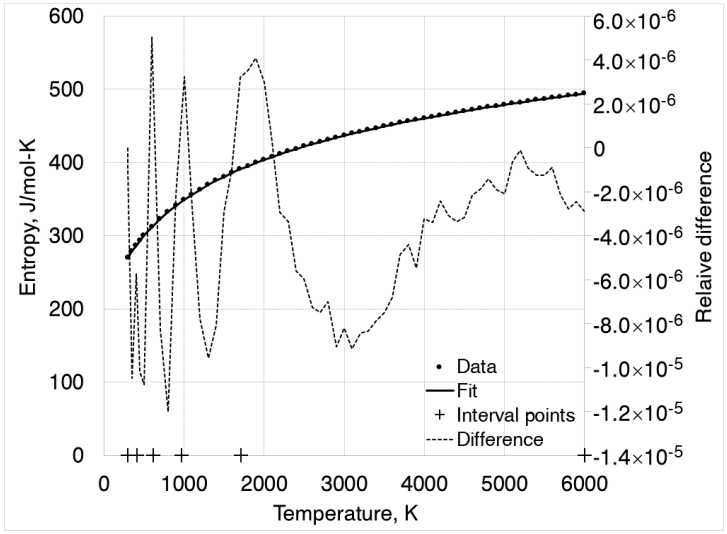


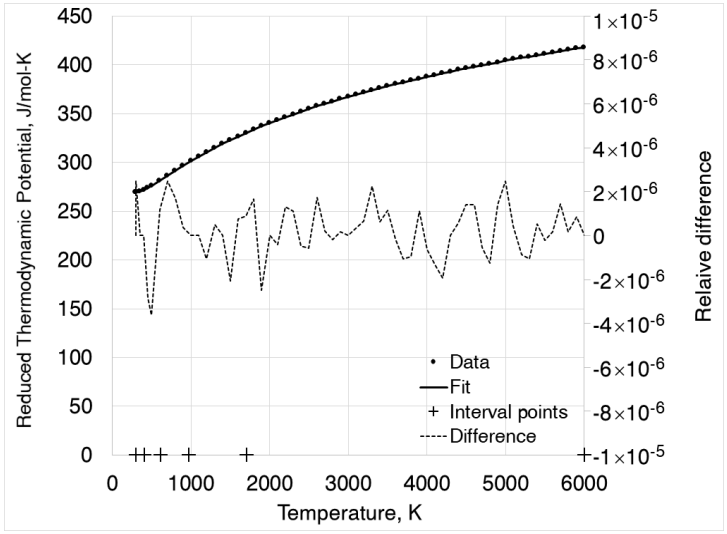
Aluminum Borate ($AlBO_2$) – Al1B1O2(g)

MW = 69.79 g/mol – $\Delta_f H^o$(298.15 K) = -541.41 kJ/mol – S^o(298.15 K) = 269.567 J/mol-K – p^o = 0.1 MPa



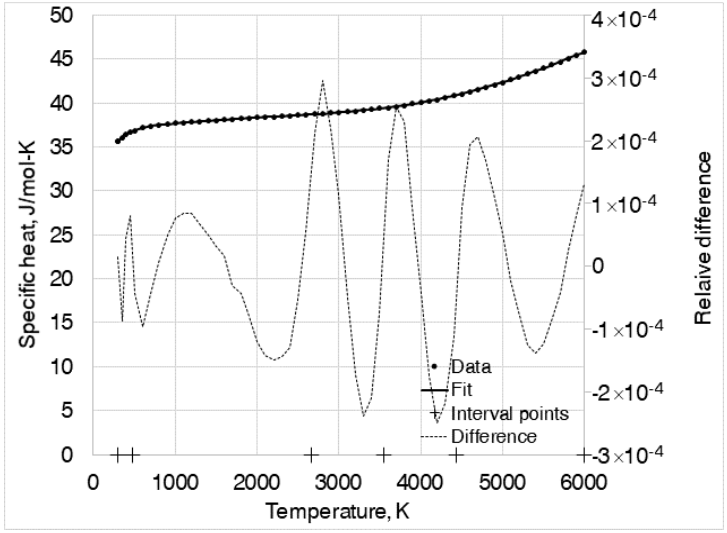


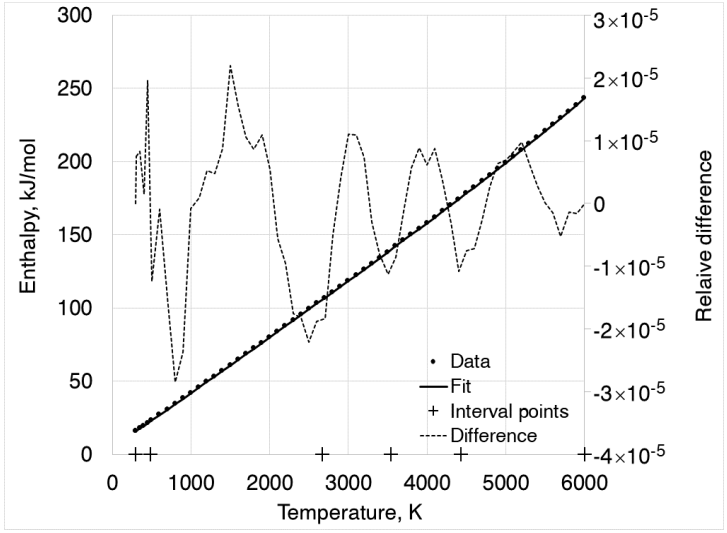


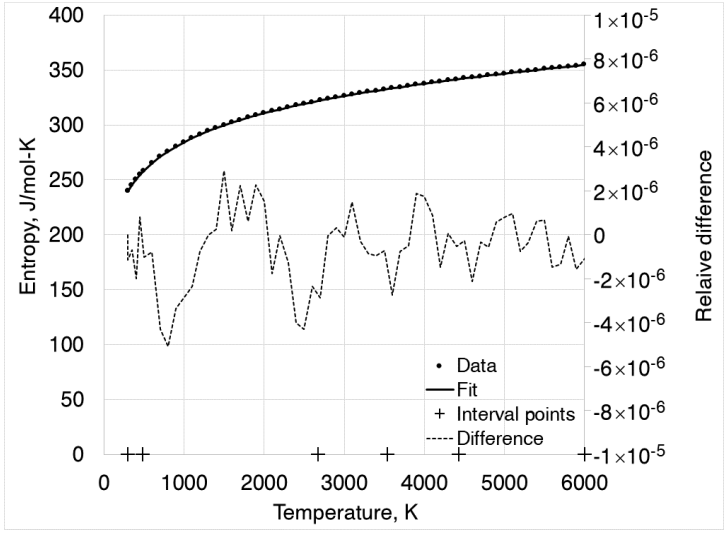


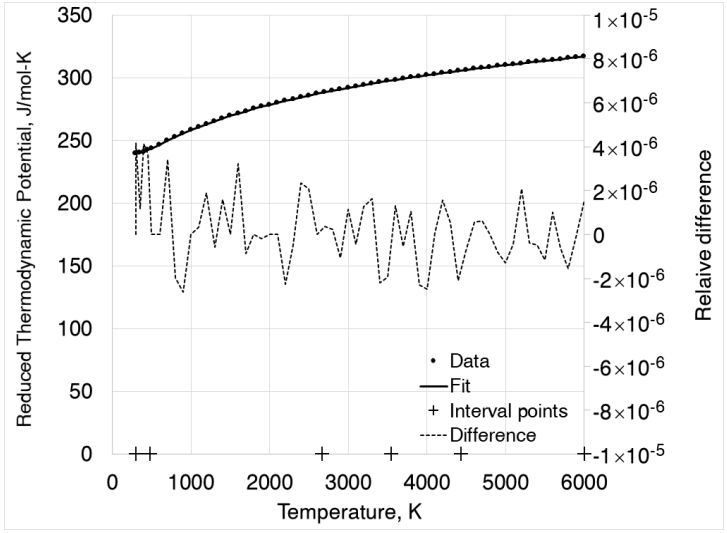
Aluminum Bromide (AlBr) – Al1Br1(g)

MW = 106.886 g/mol – $\Delta_f H^o$(298.15 K) = 15.899 kJ/mol – S^o(298.15 K) = 239.616 J/mol-K – p^o = 0.1 MPa



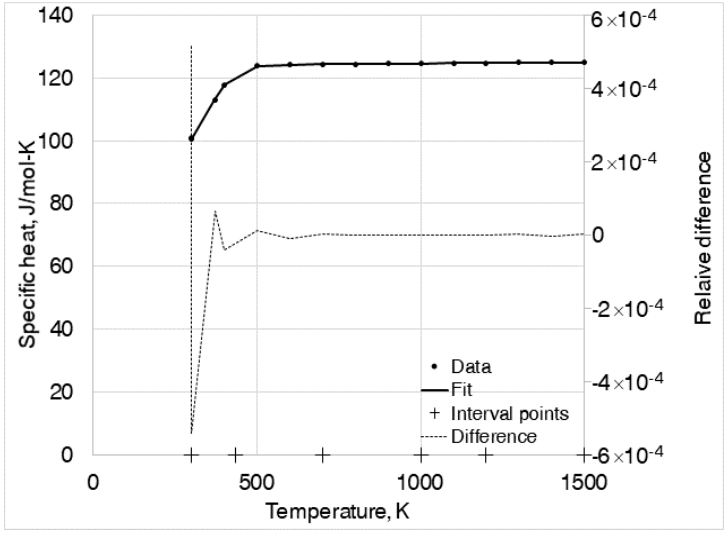


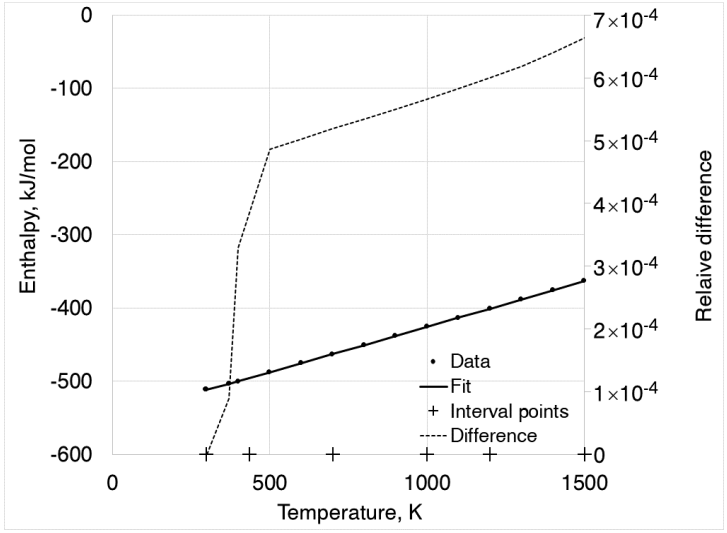


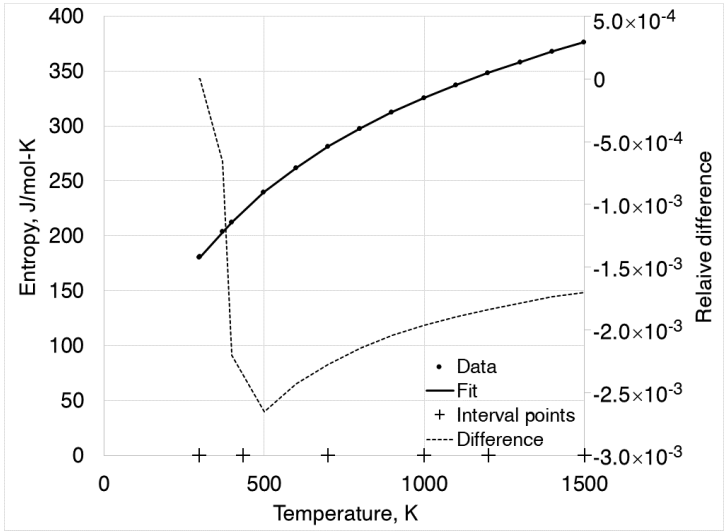


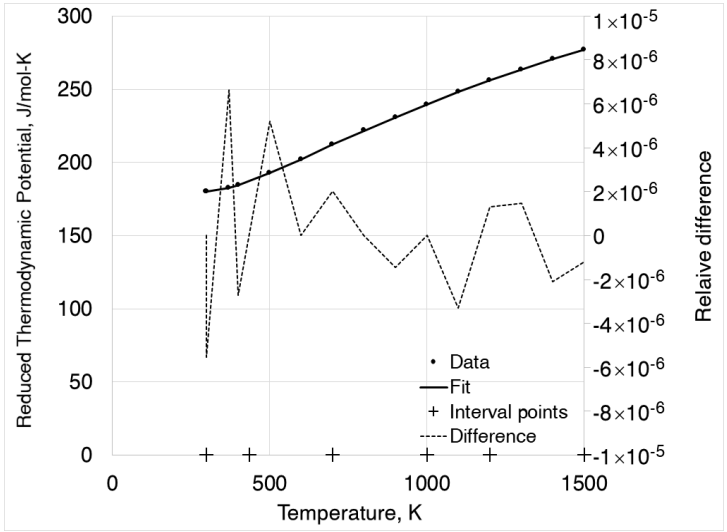
Aluminum Bromide ($AlBr_3$) – Al1Br3(cr)

MW = 266.694 g/mol – $\Delta_f H^o$(298.15 K) = -511.285 kJ/mol – S^o(298.15 K) = 180.216 J/mol-K – p^o = 0.1 MPa



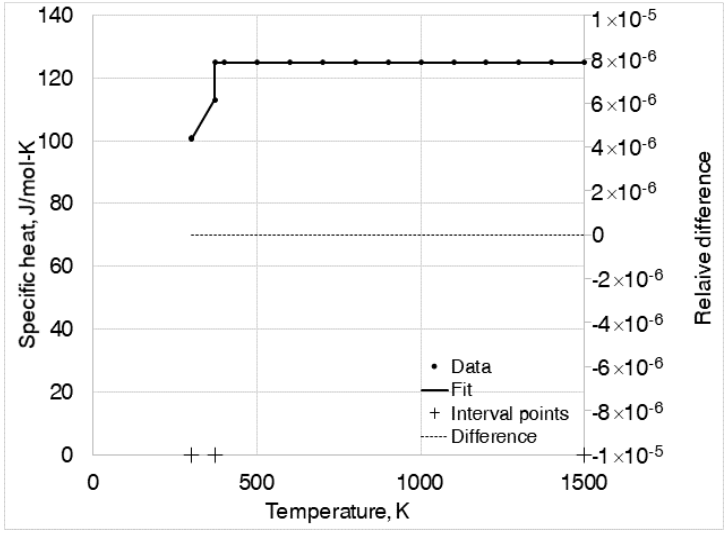


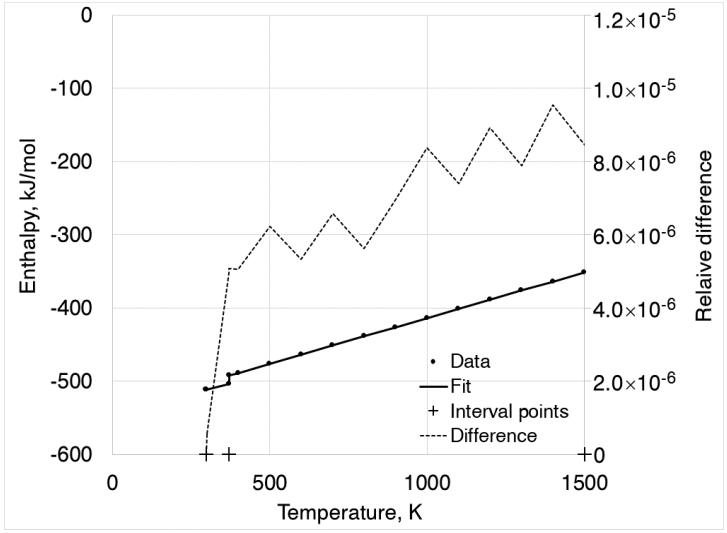


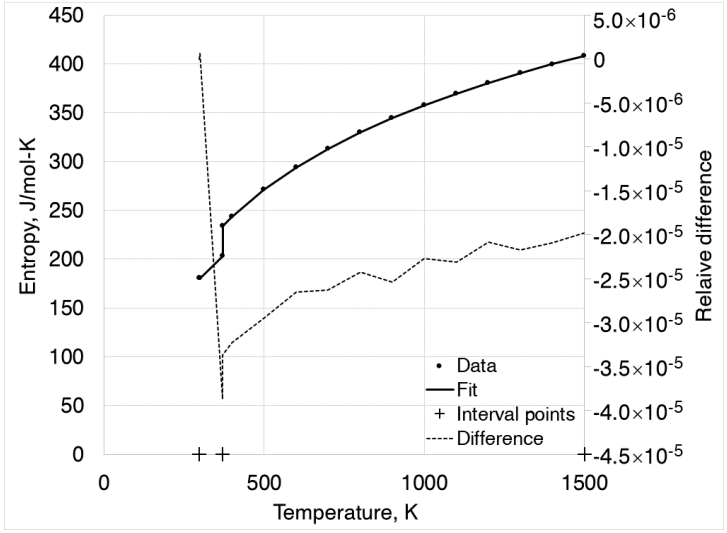


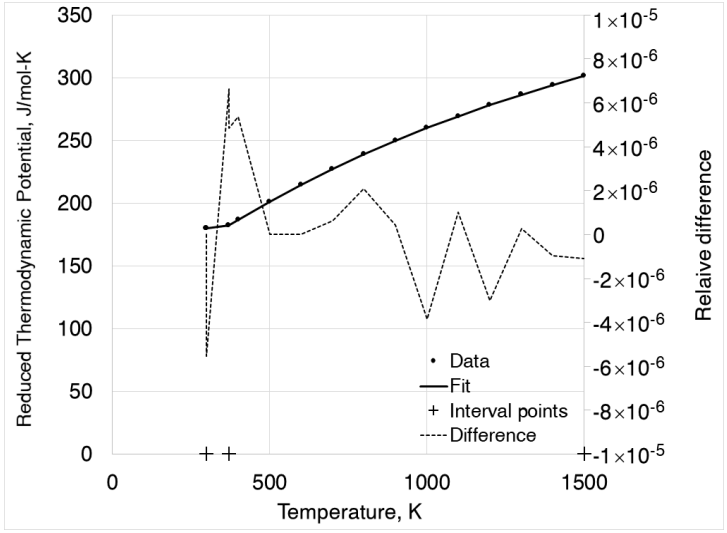
Aluminum Bromide ($AlBr_3$) – Al1Br3(cr,l)

MW = 266.694 g/mol – $\Delta_f H^o$(298.15 K) = -511.285 kJ/mol – S^o(298.15 K) = 180.216 J/mol-K – p^o = 0.1 MPa



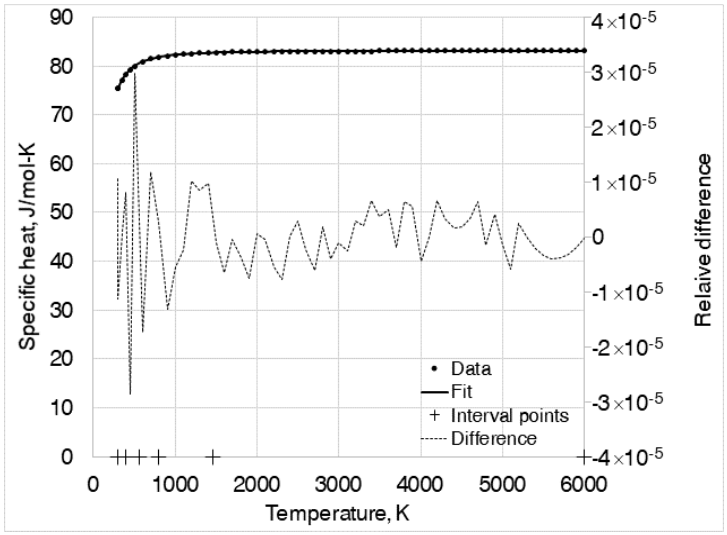


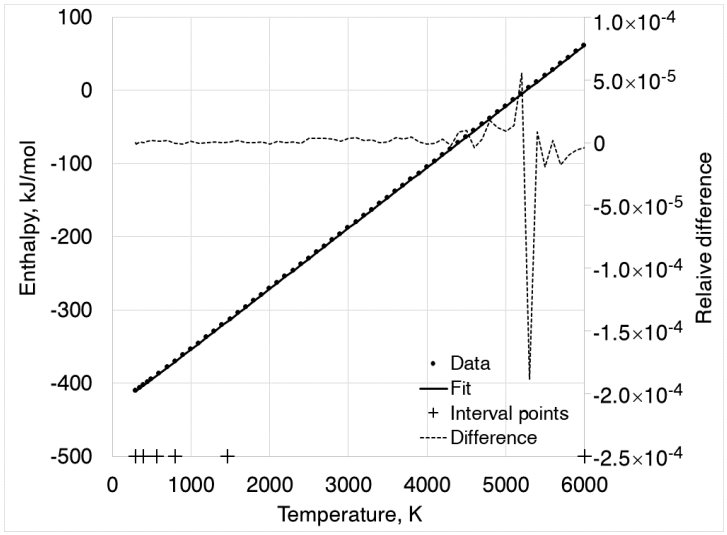


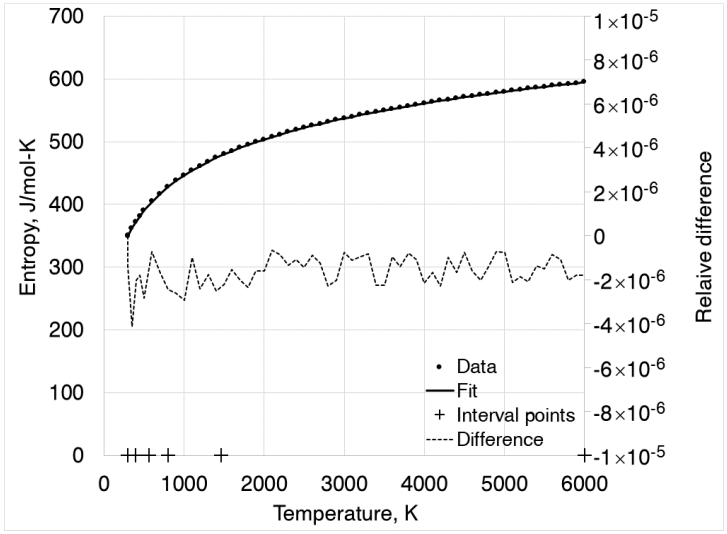


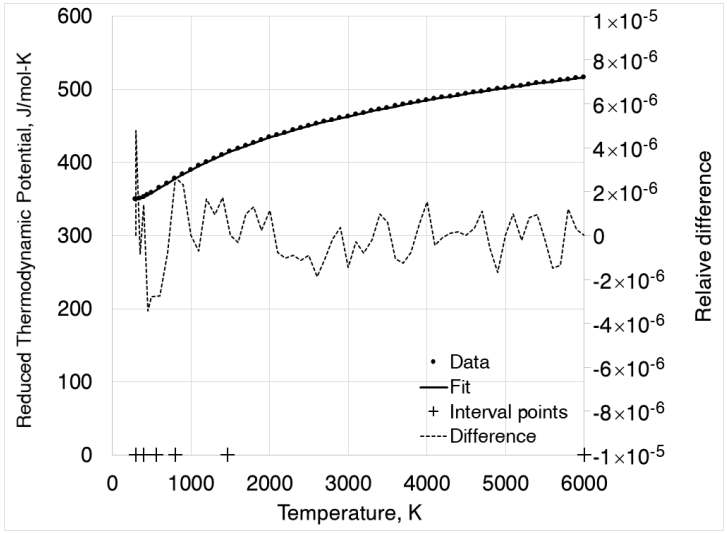
Aluminum Bromide ($AlBr_3$) – Al1Br3(g)

MW = 266.694 g/mol – $\Delta_f H^o$(298.15 K) = -410.45 kJ/mol – S^o(298.15 K) = 349.441 J/mol-K – p^o = 0.1 MPa



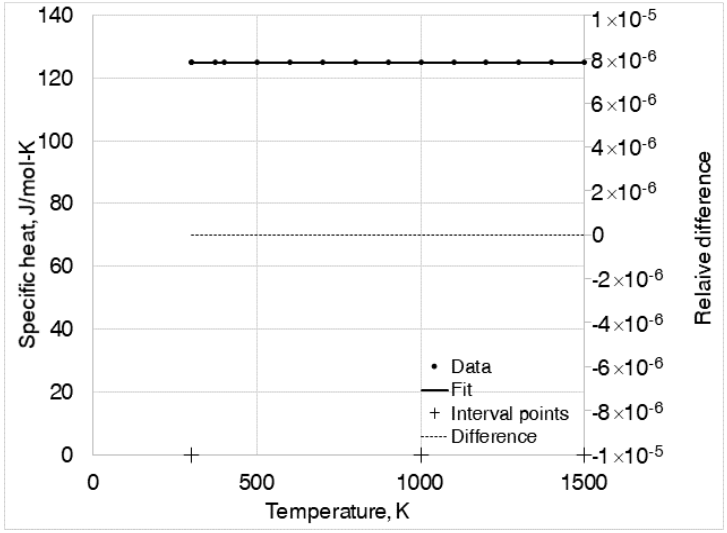


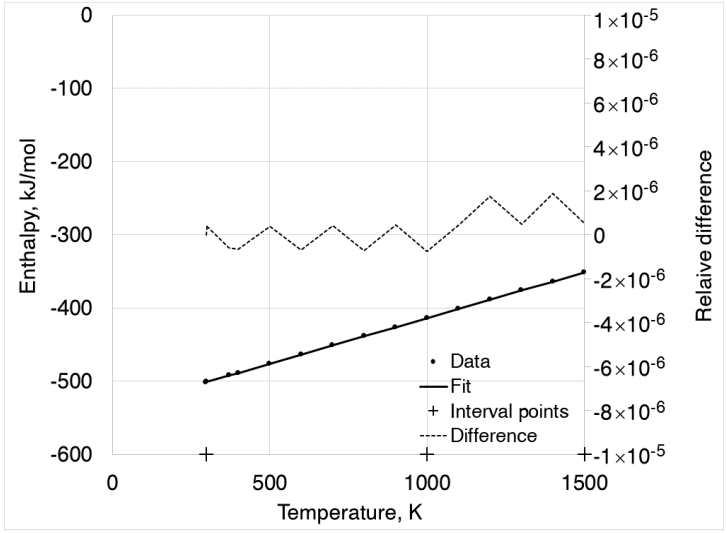


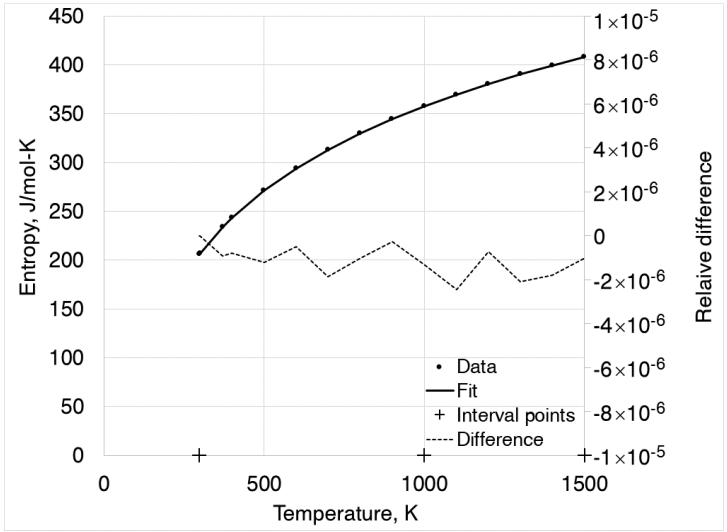


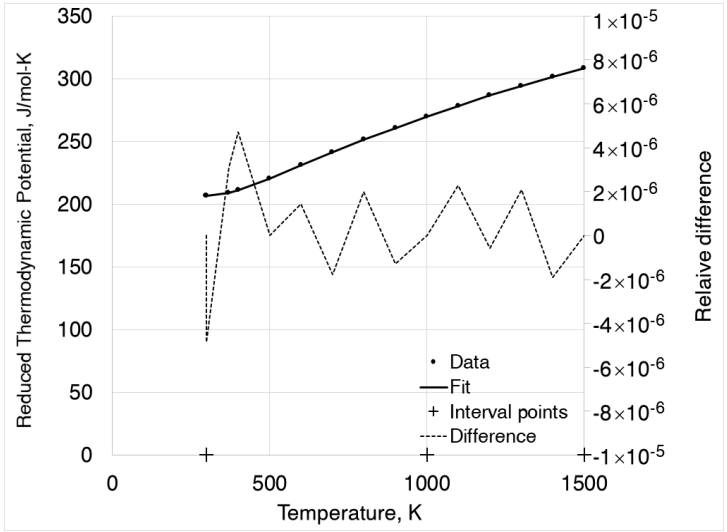
Aluminum Bromide ($AlBr_3$) – Al1Br3(l)

MW = 266.694 g/mol – $\Delta_f H^o$(298.15 K) = -501.381 kJ/mol – S^o(298.15 K) = 206.481 J/mol-K – p^o = 0.1 MPa



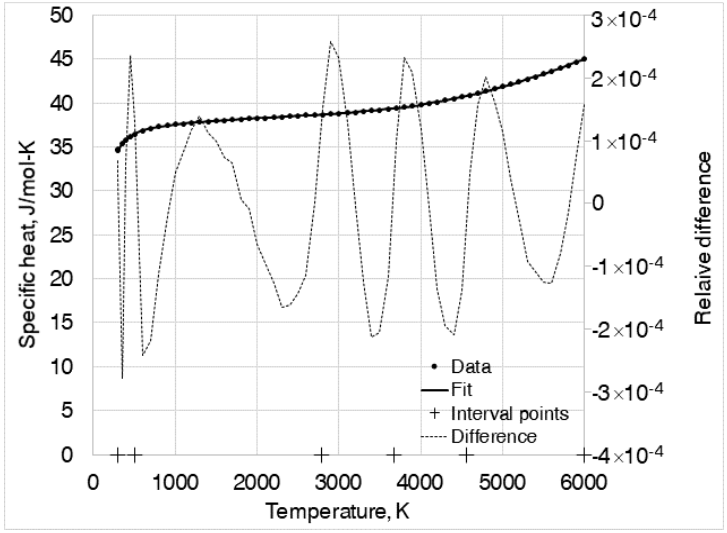


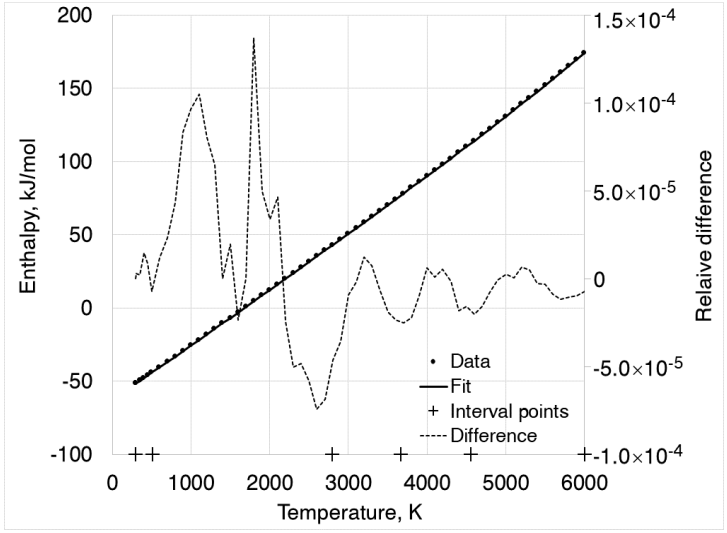


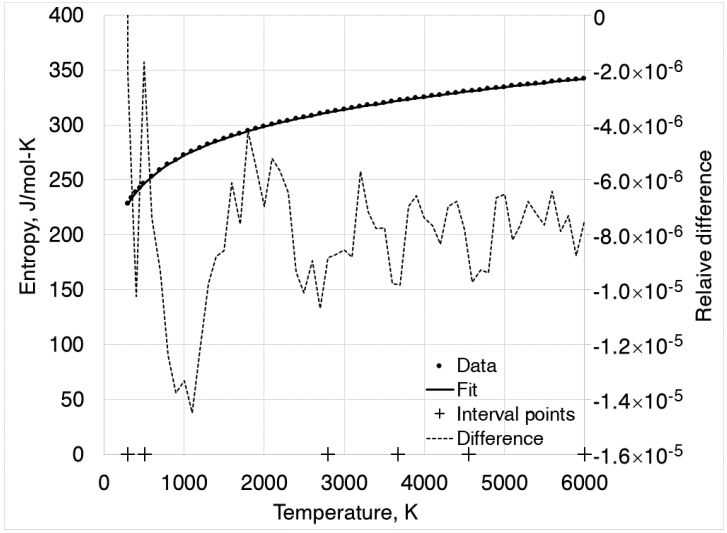


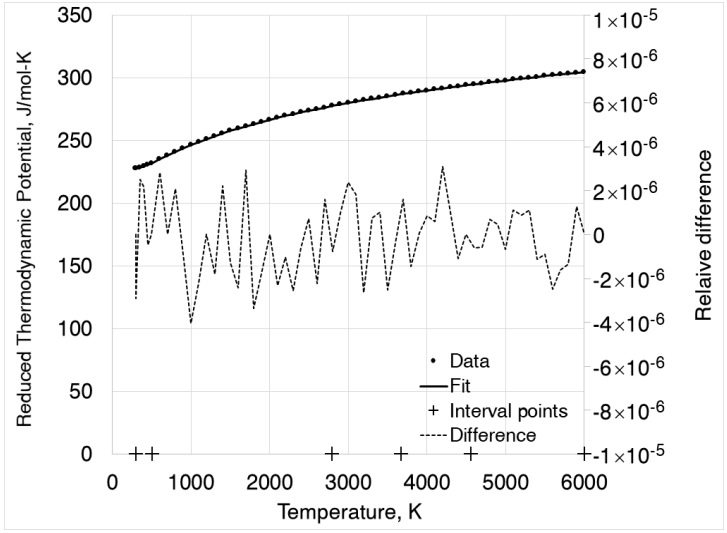
Aluminum Chloride (AlCl) – Al1Cl1(g)

MW = 62.435 g/mol – $\Delta_f H^o$(298.15 K) = -51.463 kJ/mol – S^o(298.15 K) = 227.957 J/mol-K – p^o = 0.1 MPa



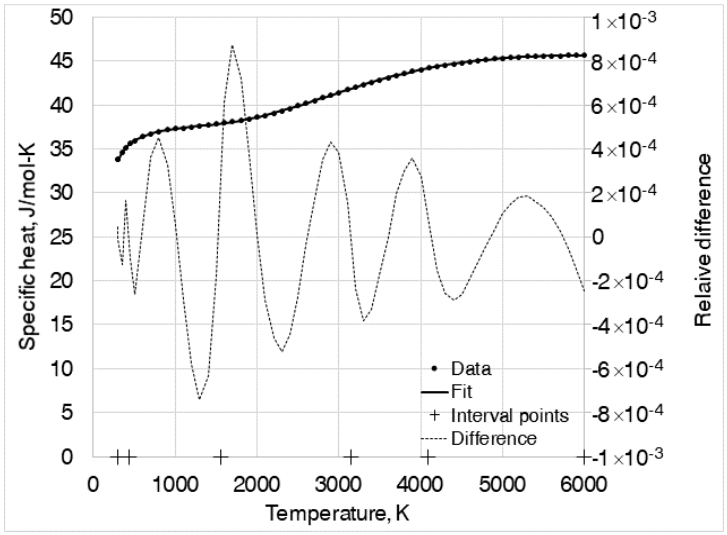


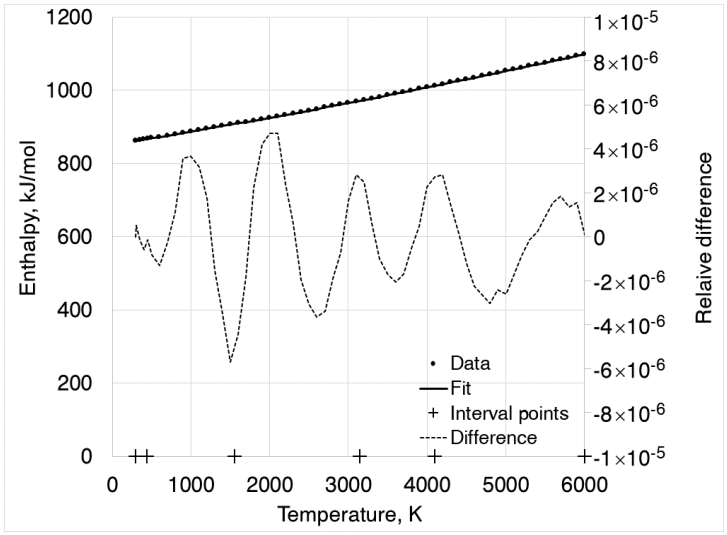


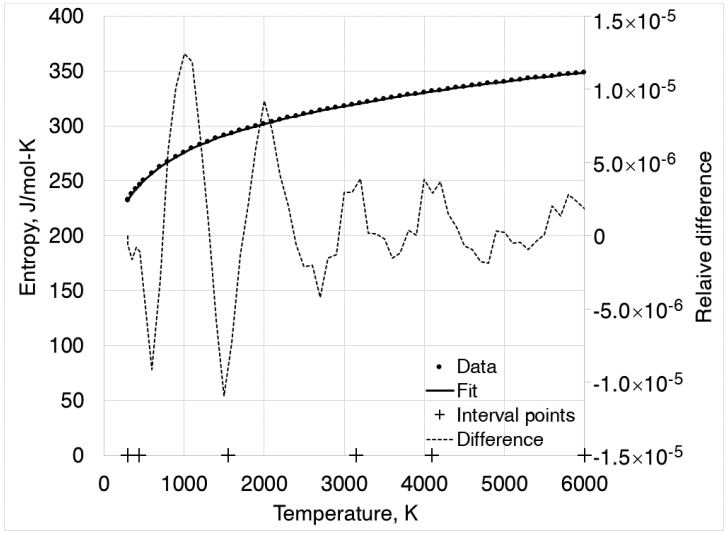


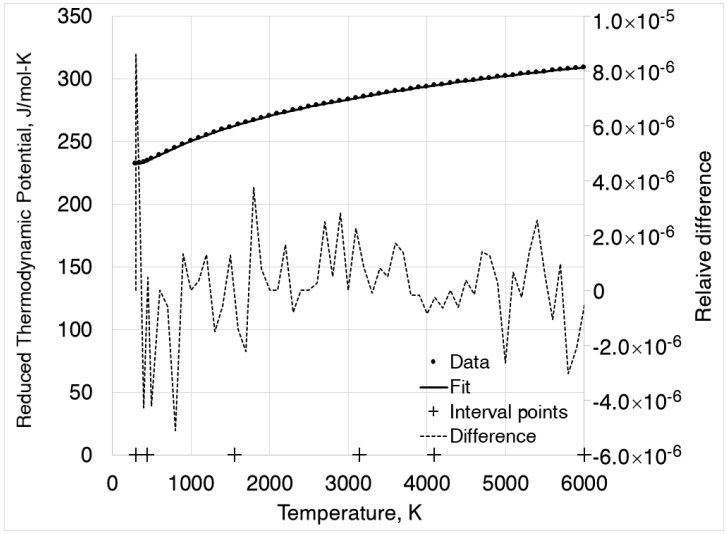
Aluminum Chloride, Ion ($AlCl^+$) – Al1Cl1+(g)

MW = 62.434 g/mol – $\Delta_f H^o$(298.15 K) = 861.904 kJ/mol – S^o(298.15 K) = 232.187 J/mol-K – p^o = 0.1 MPa



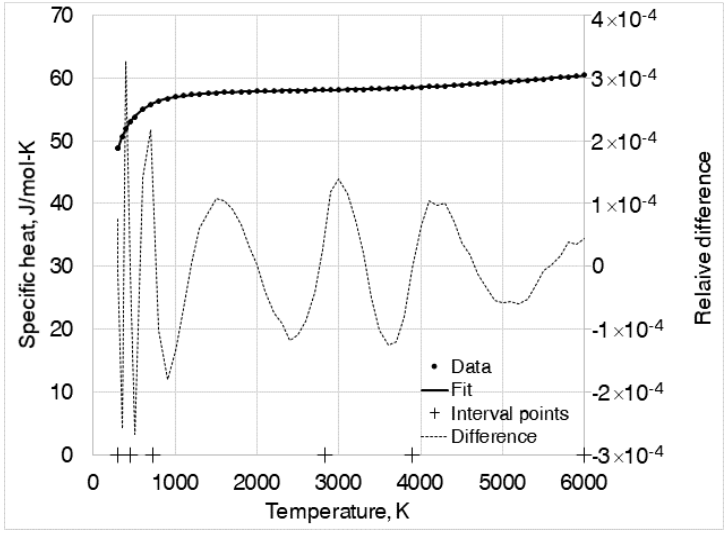


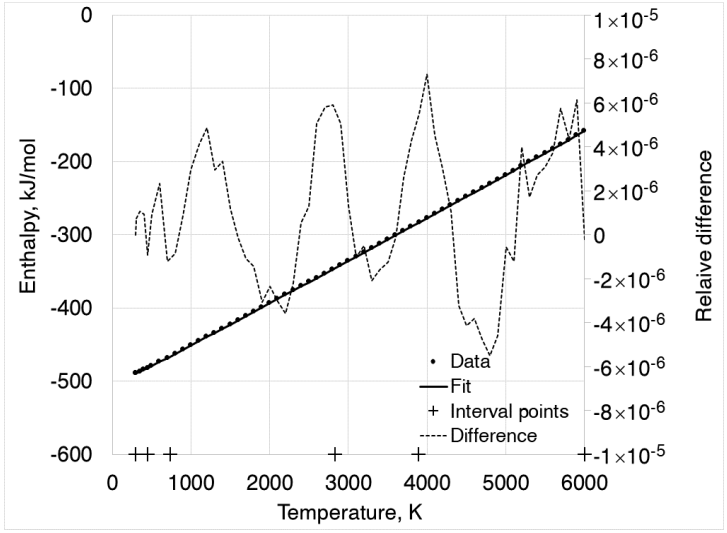


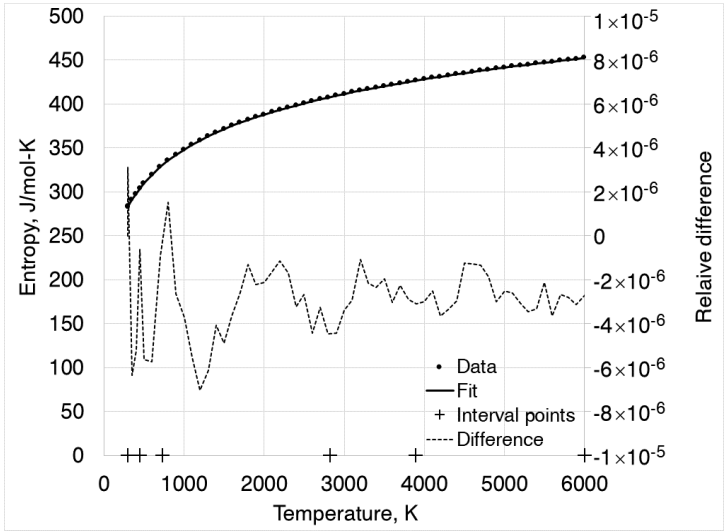


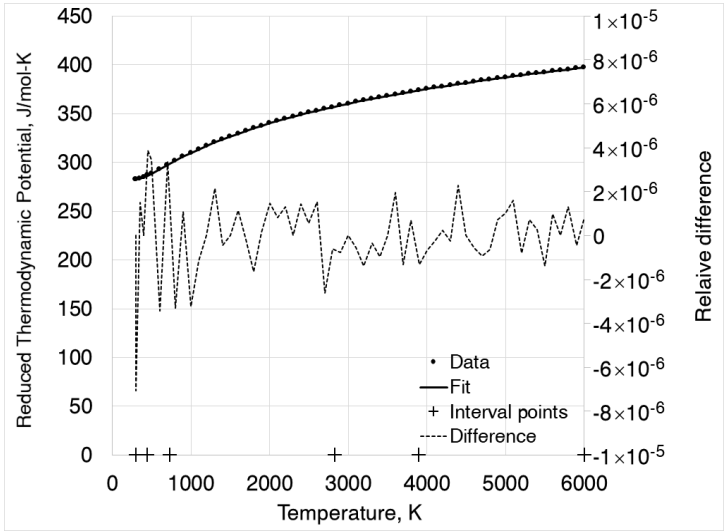
Aluminum Chloride Fluoride (AlClF) – Al1Cl1F1(g)

MW = 81.433 g/mol – $\Delta_f H^o$(298.15 K) = -489.528 kJ/mol – S^o(298.15 K) = 282.895 J/mol-K – p^o = 0.1 MPa



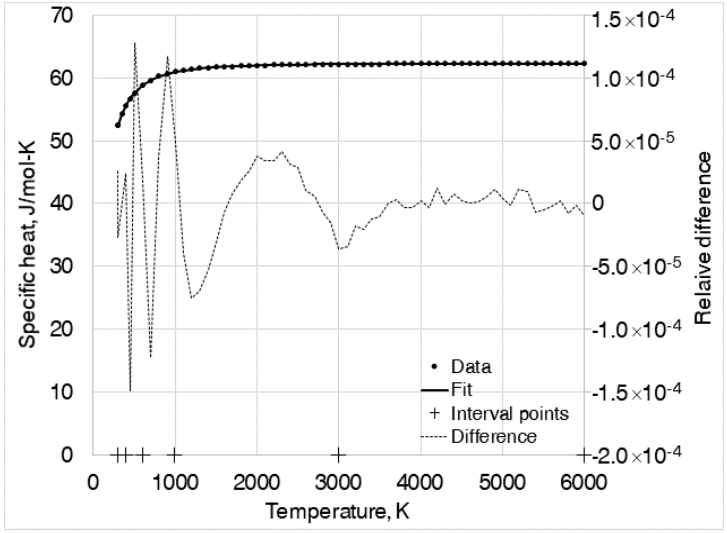


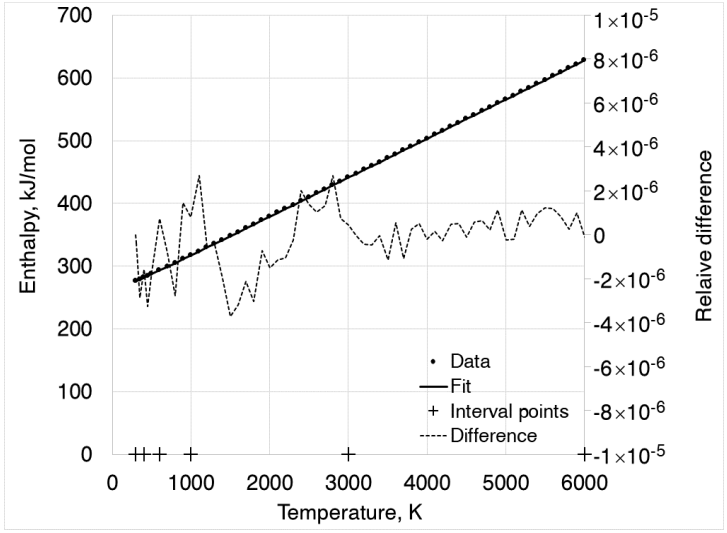


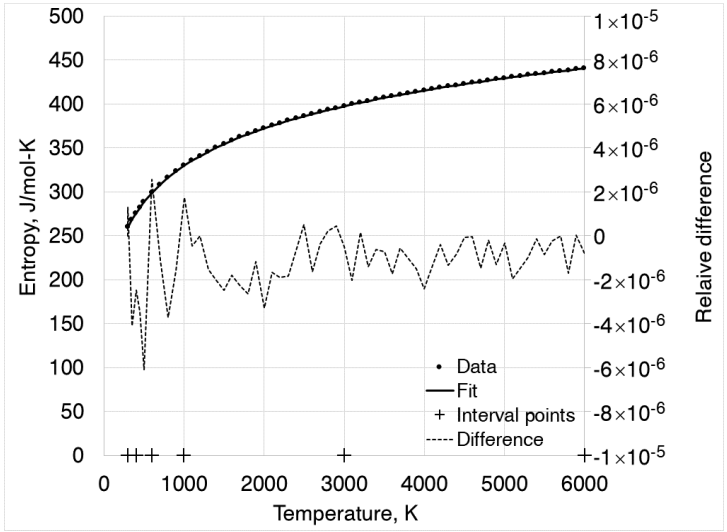


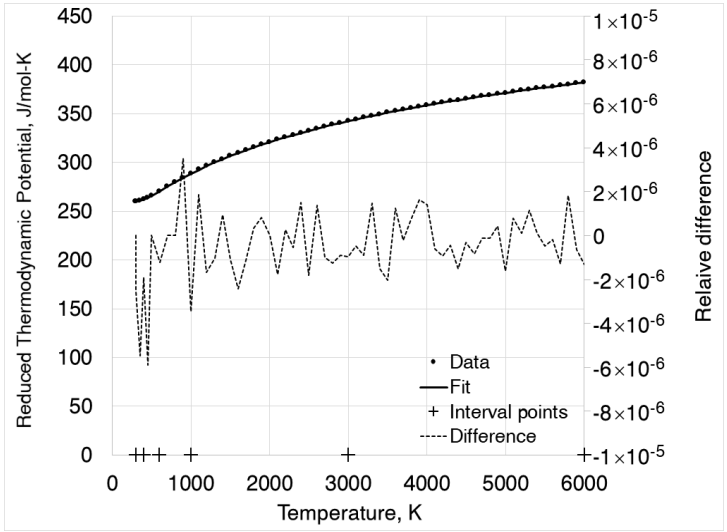
Aluminum Chloride Fluoride, Ion ($AlClF^+$) – Al1Cl1F1+(g)

MW = 81.432 g/mol – $\Delta_f H^o$(298.15 K) = 276.144 kJ/mol – S^o(298.15 K) = 259.722 J/mol-K – p^o = 0.1 MPa



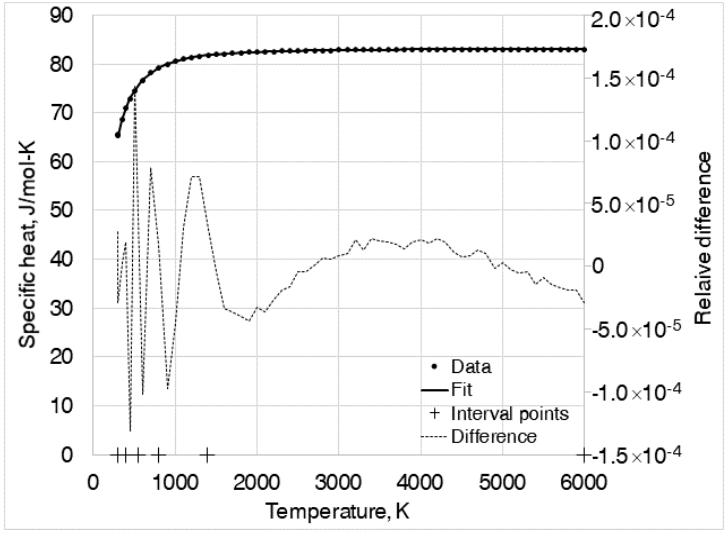


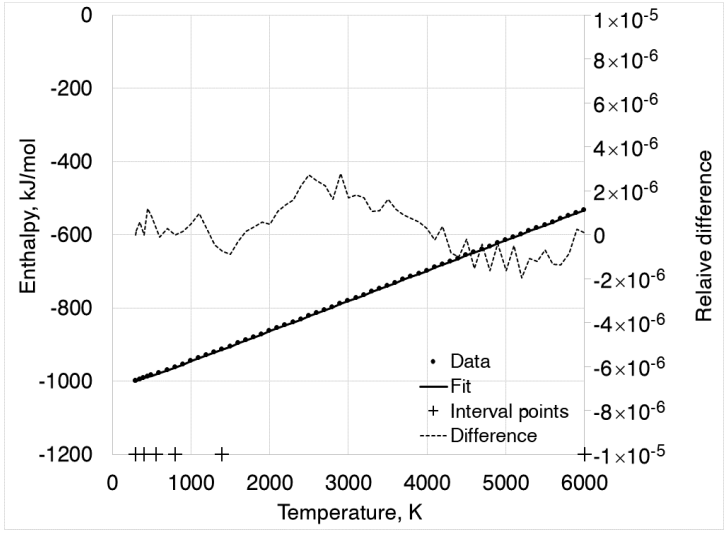


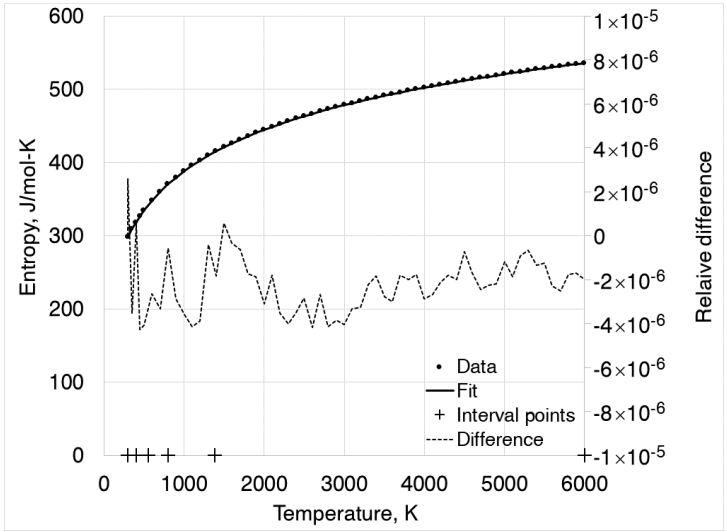


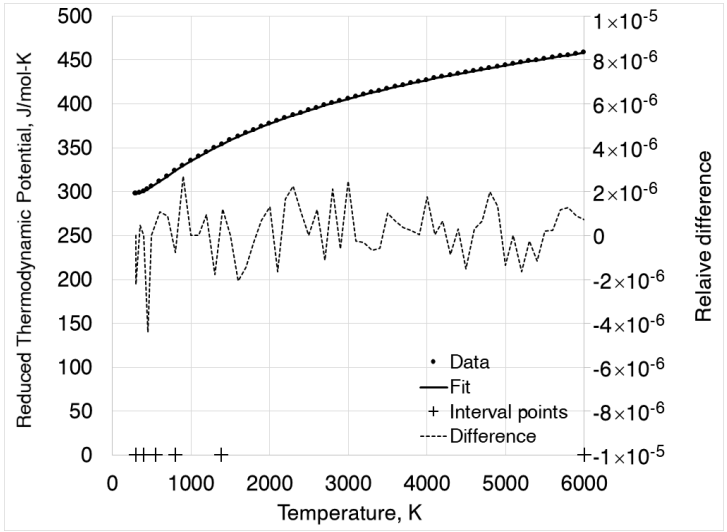
Aluminum Chloride Fluoride ($AlClF_2$) – Al1Cl1F2(g)

MW = 100.431 g/mol – $\Delta_f H^o$(298.15 K) = -999.139 kJ/mol – S^o(298.15 K) = 297.839 J/mol-K – p^o = 0.1 MPa



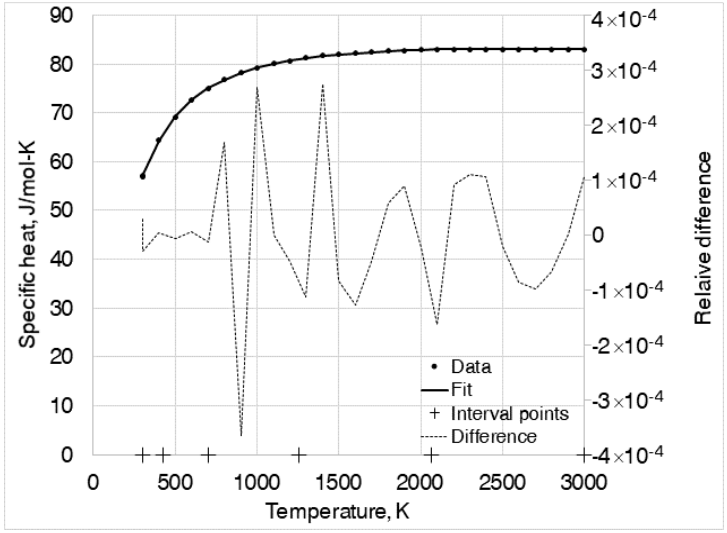


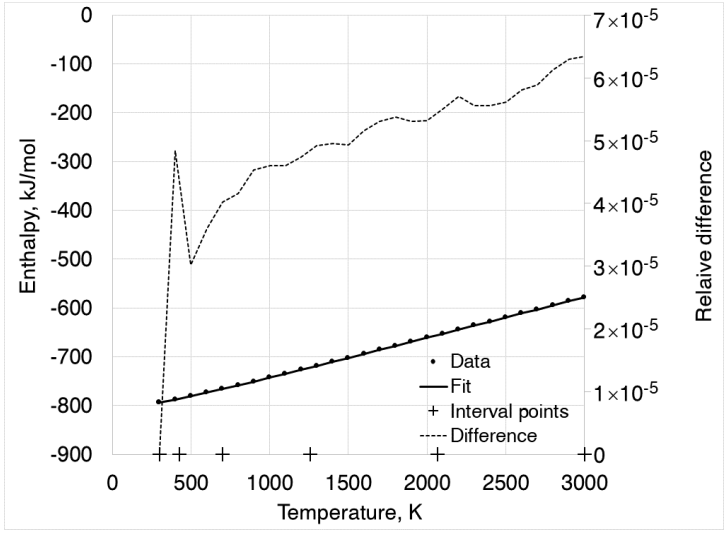


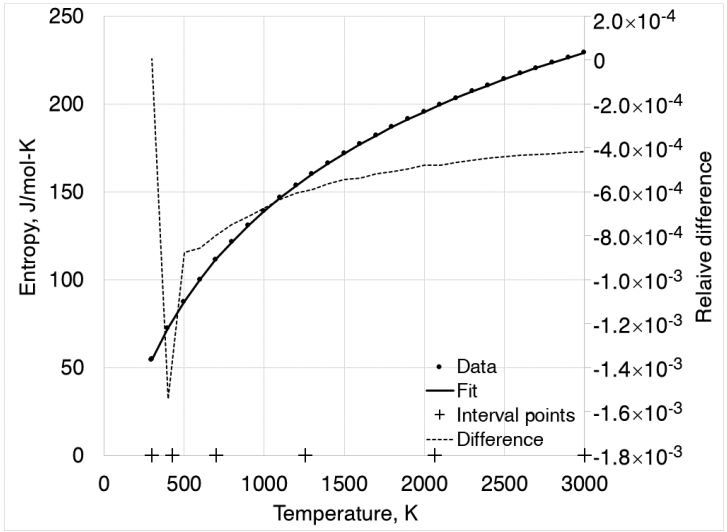


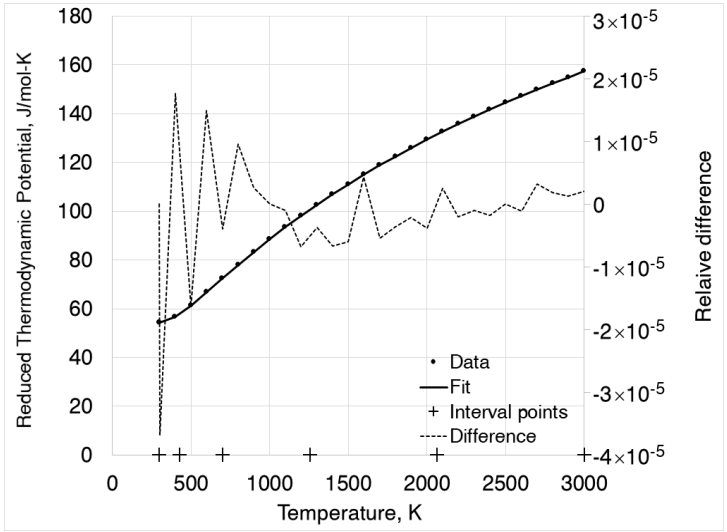
Aluminum Chloride Oxide (OAlCl) – Al1Cl1O1(cr)

MW = 78.434 g/mol – $\Delta_f H^o$(298.15 K) = -793.286 kJ/mol – S^o(298.15 K) = 54.392 J/mol-K – p^o = 0.1 MPa



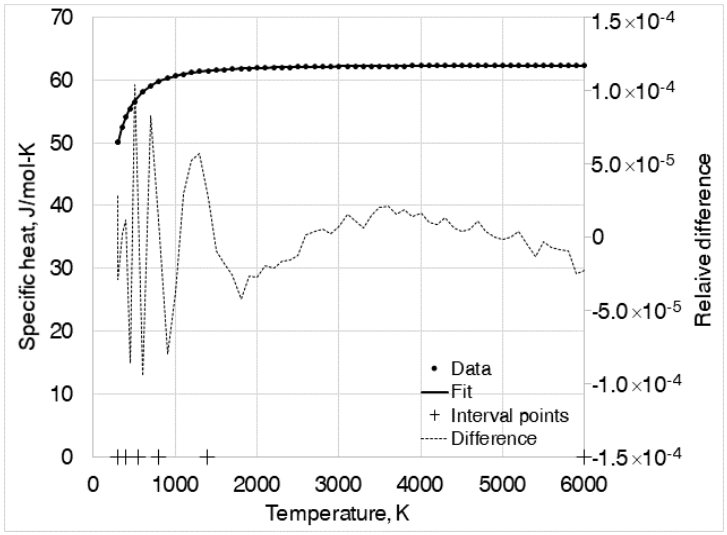


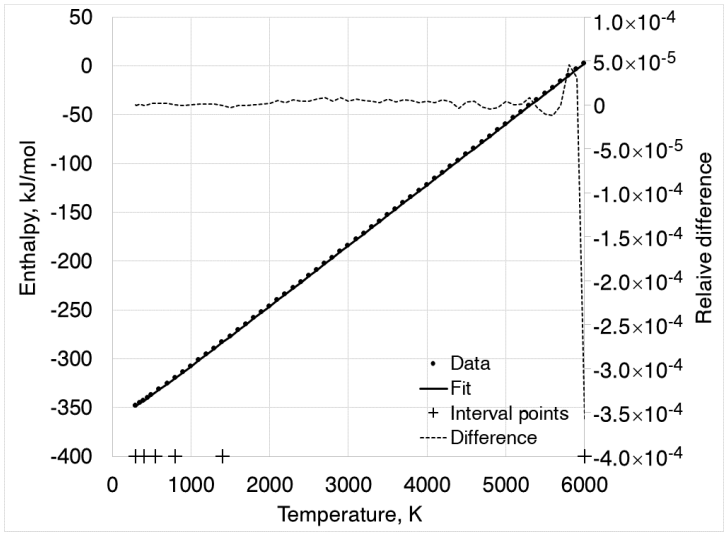


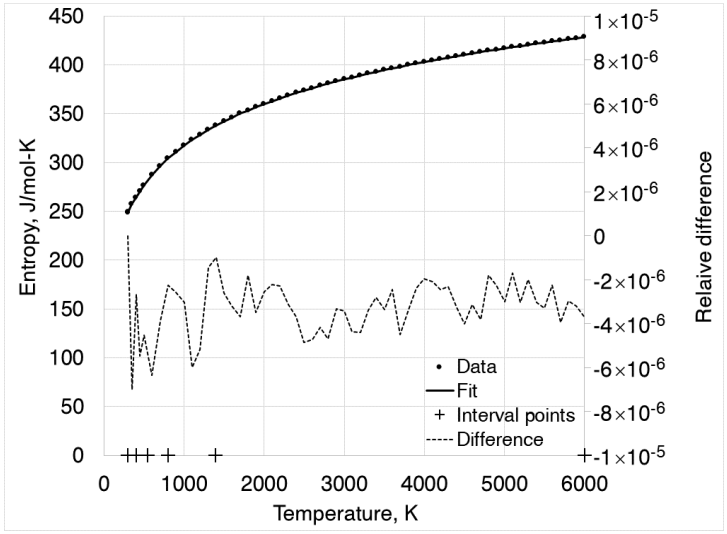


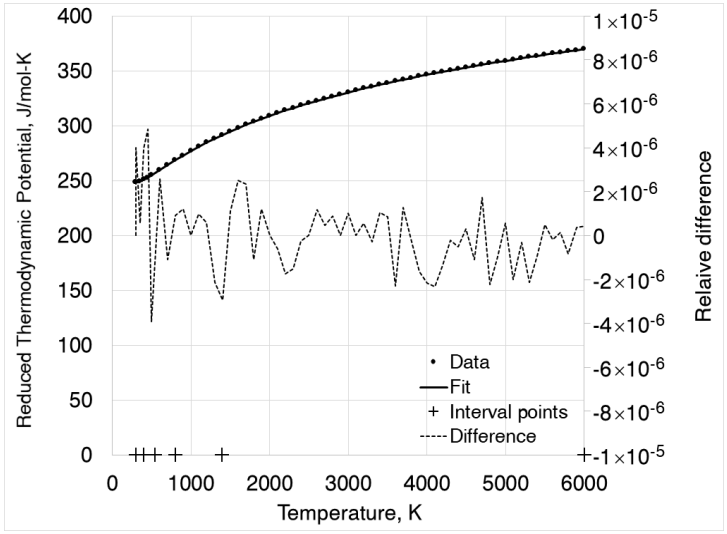
Aluminum Chloride Oxide (OAlCl) – Al1Cl1O1(g)

MW = 78.434 g/mol – $\Delta_f H^o$(298.15 K) = -348.109 kJ/mol – S^o(298.15 K) = 248.931 J/mol-K – p^o = 0.1 MPa



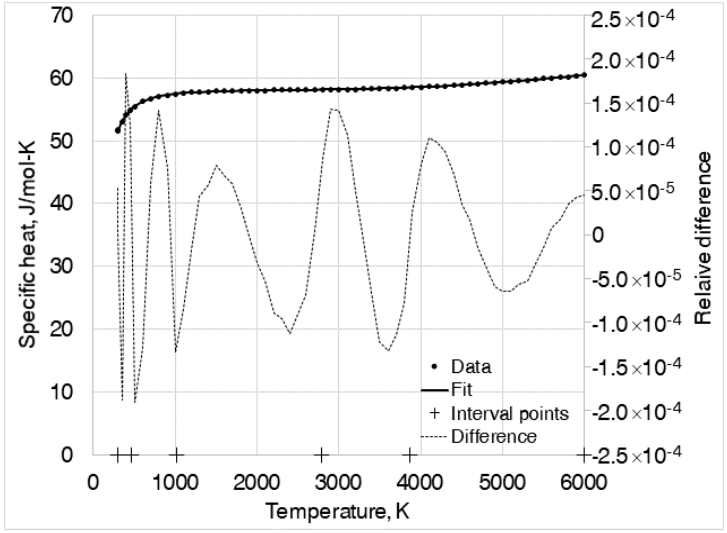


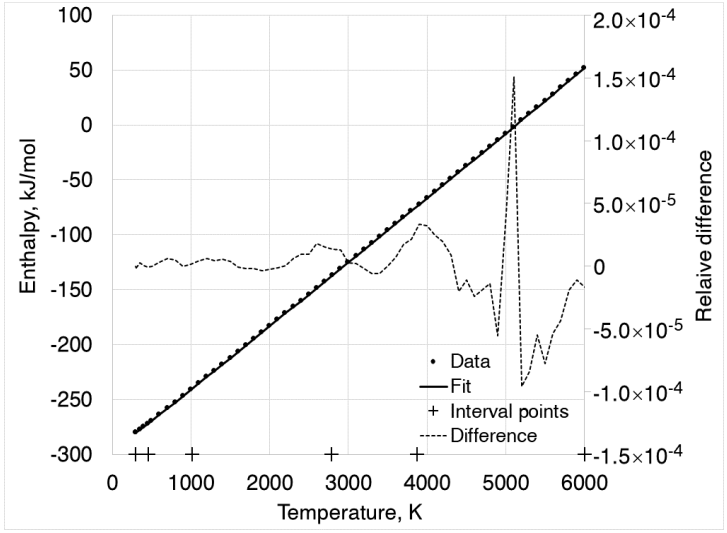


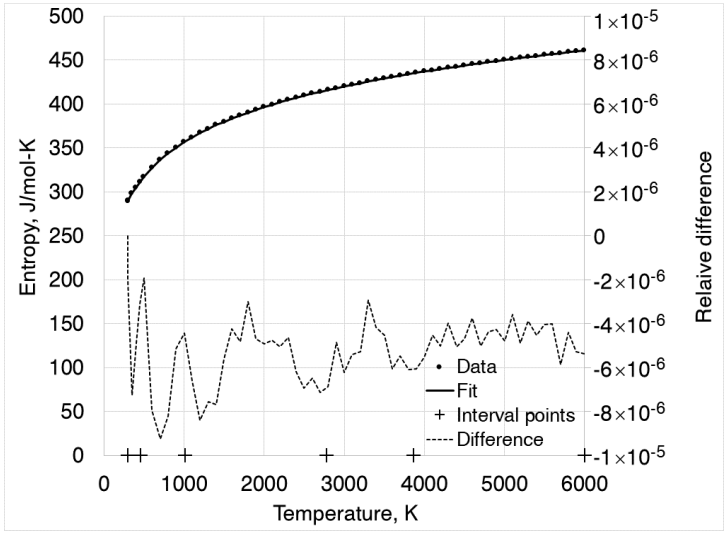


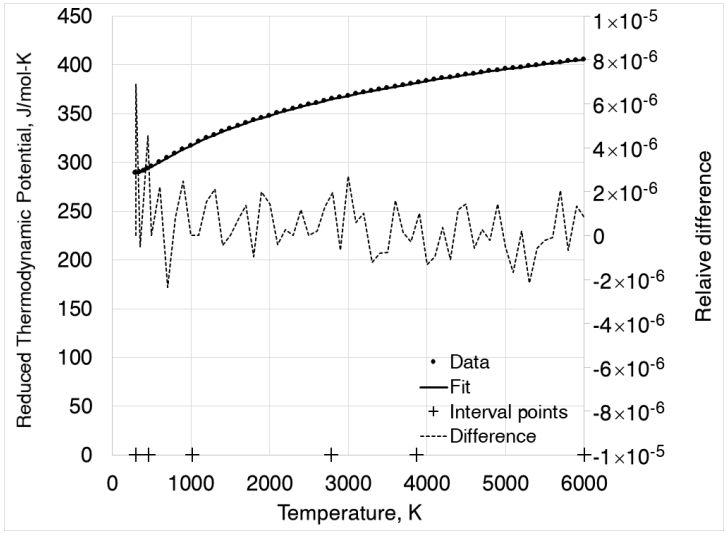
Aluminum Chloride ($AlCl_2$) – Al1Cl2(g)

MW = 97.888 g/mol – $\Delta_f H^o$(298.15 K) = -280.328 kJ/mol – S^o(298.15 K) = 289.42 J/mol-K – p^o = 0.1 MPa



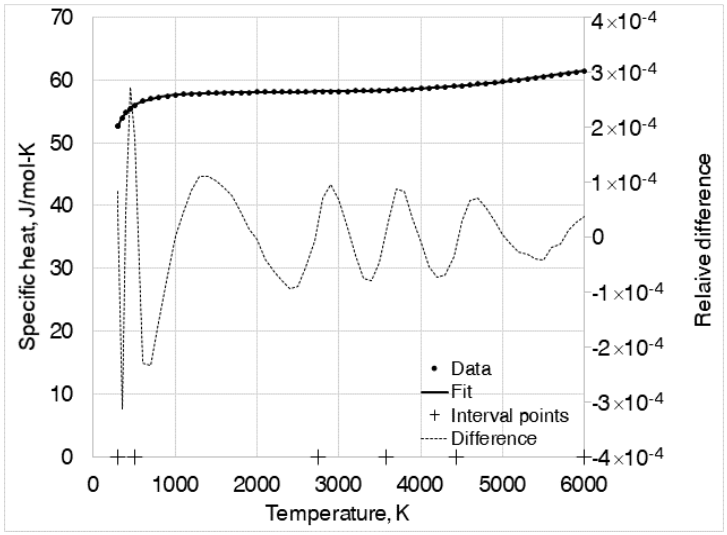


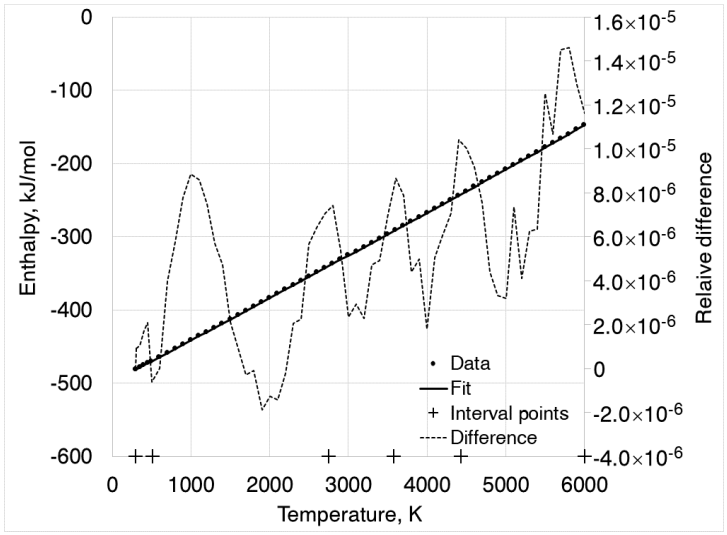


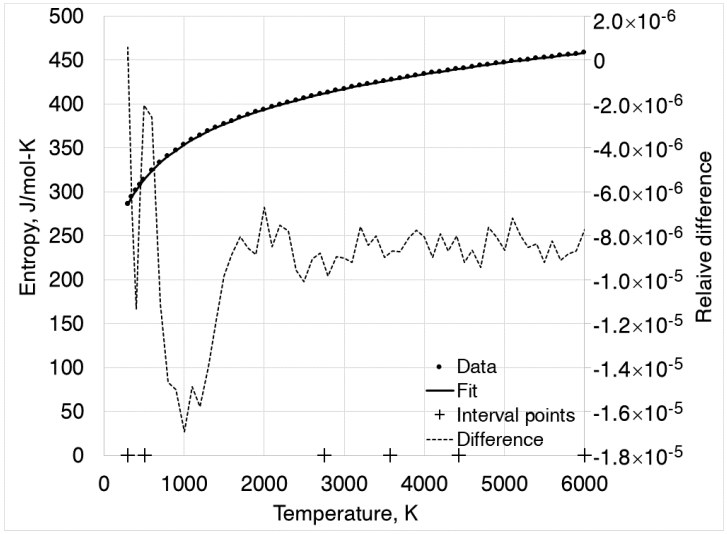


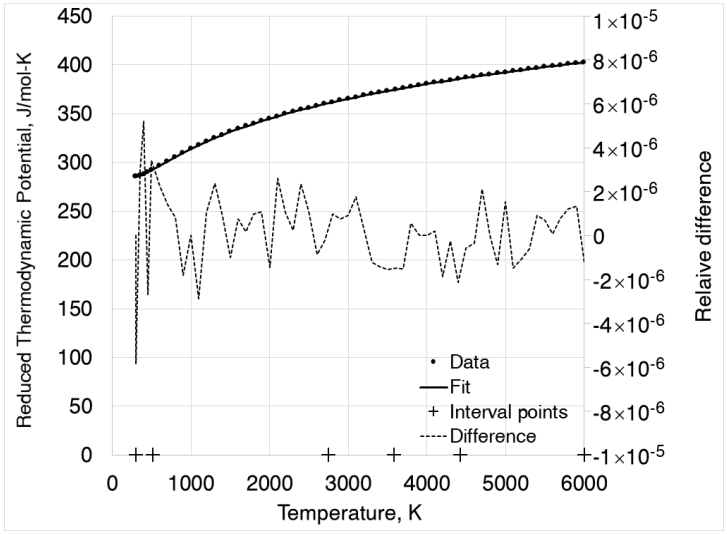
Aluminum Chloride, Ion ($AlCl_2^-$) – Al1Cl2-(g)

MW = 97.888 g/mol – $\Delta_f H^o$(298.15 K) = -481.16 kJ/mol – S^o(298.15 K) = 285.778 J/mol-K – p^o = 0.1 MPa



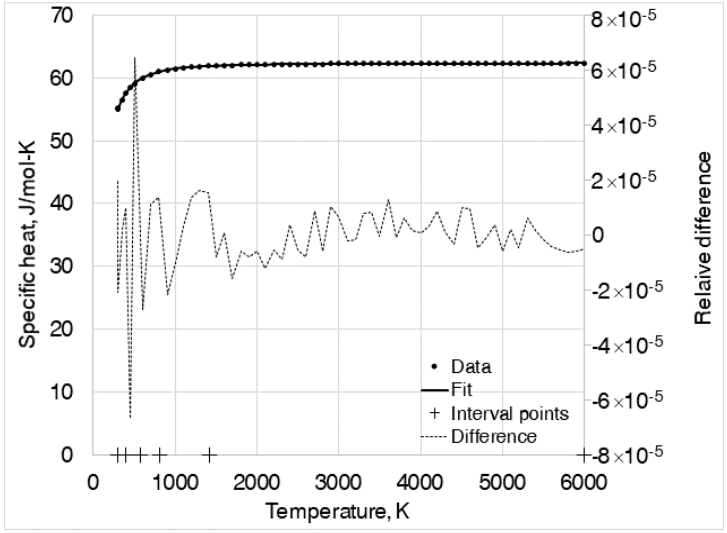


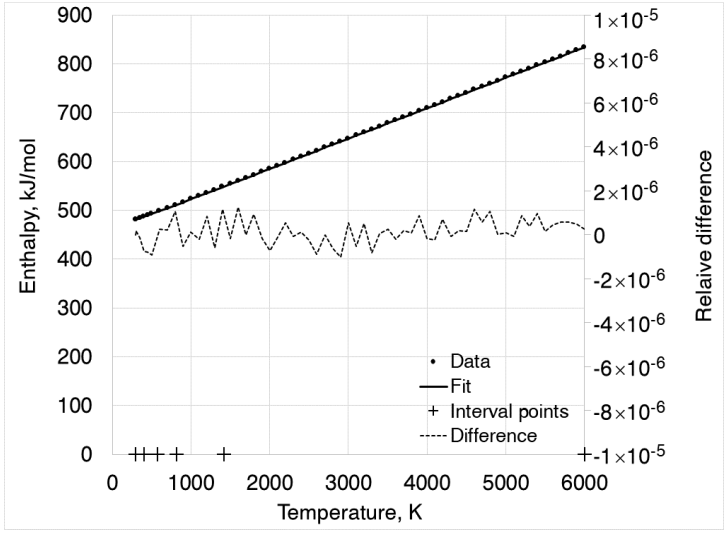


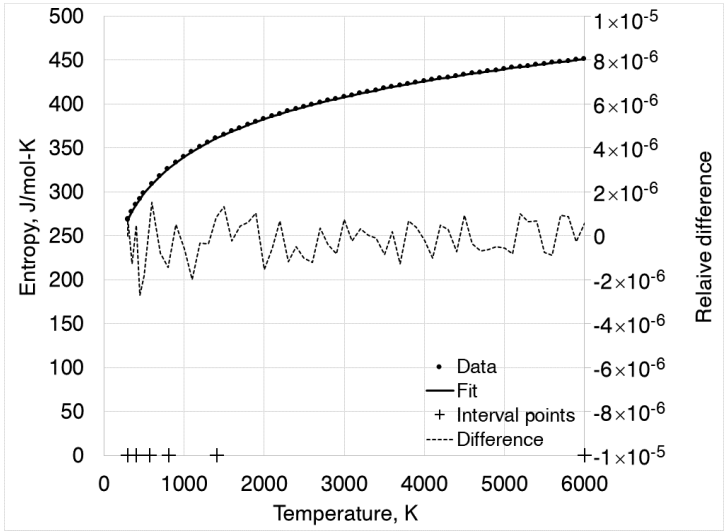


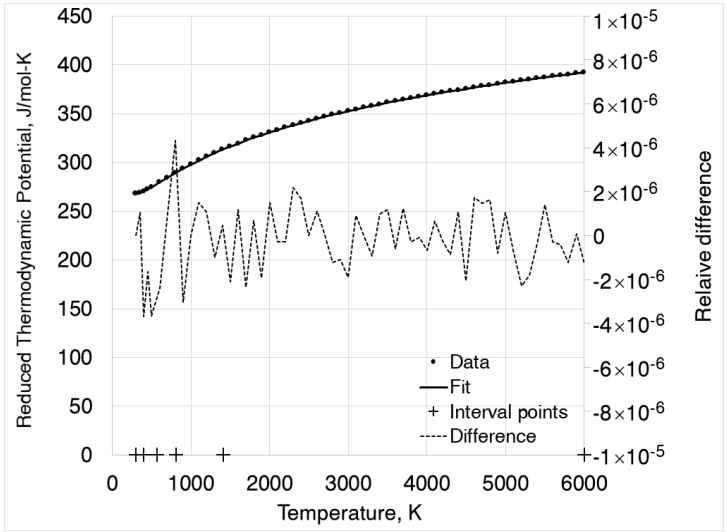
Aluminum Chloride, Ion ($AlCl_2^+$) – Al1Cl2+(g)

MW = 97.887 g/mol – $\Delta_f H^o$(298.15 K) = 481.16 kJ/mol – S^o(298.15 K) = 268.261 J/mol-K – p^o = 0.1 MPa



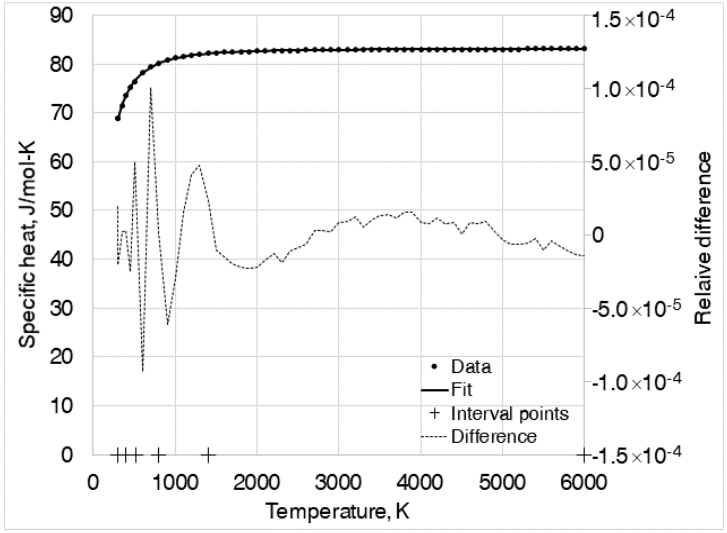


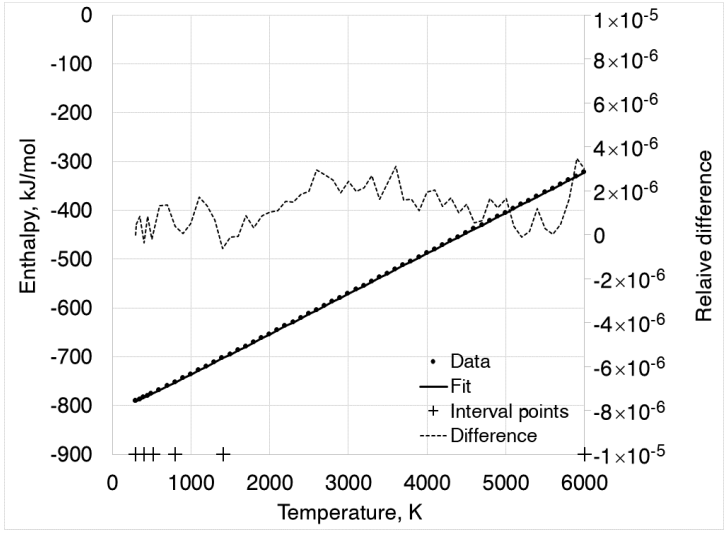


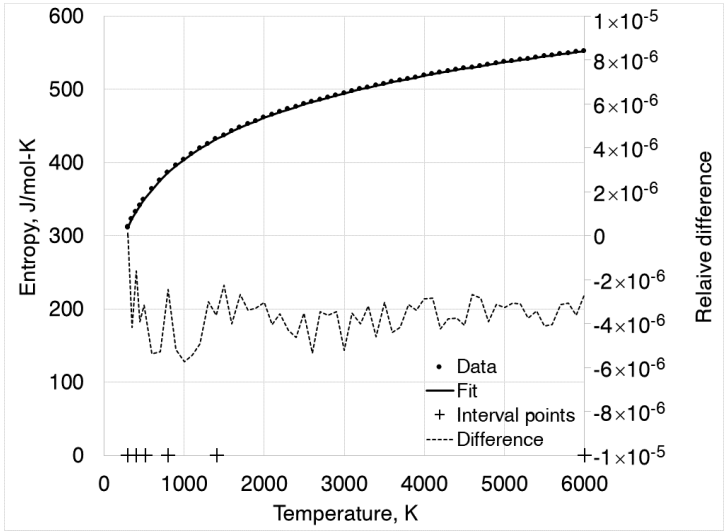


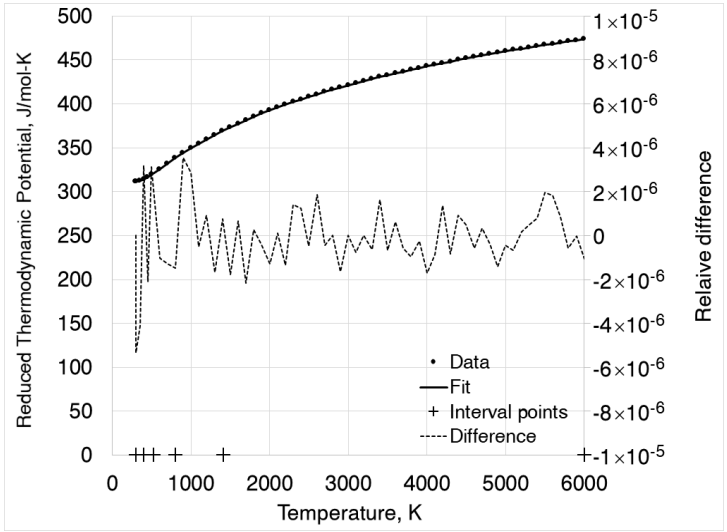
Aluminum Chloride Fluoride ($AlCl_2F$) – Al1Cl2F1(g)

MW = 116.886 g/mol – $\Delta_f H^o$(298.15 K) = -790.776 kJ/mol – S^o(298.15 K) = 311.403 J/mol-K – p^o = 0.1 MPa



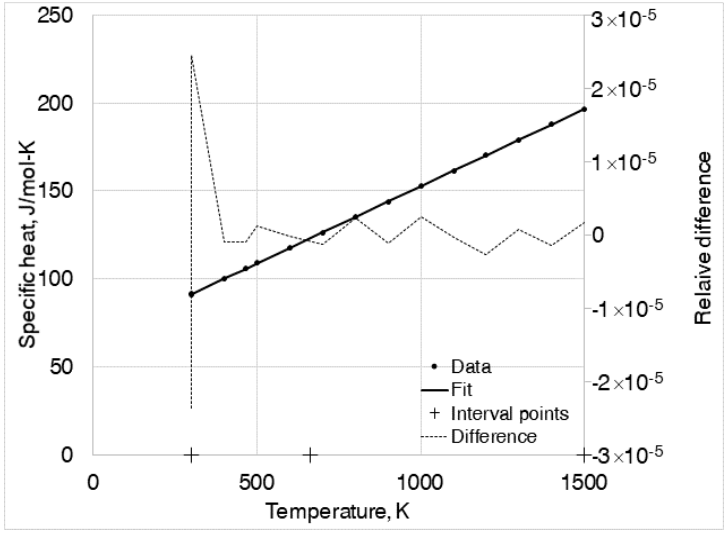


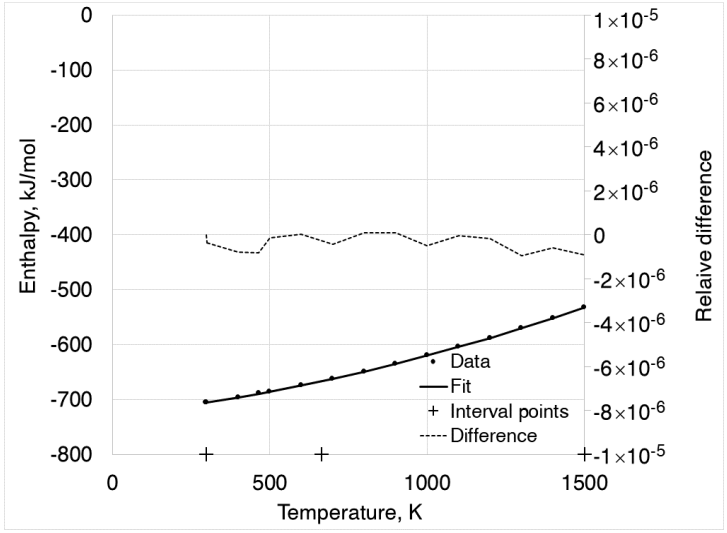


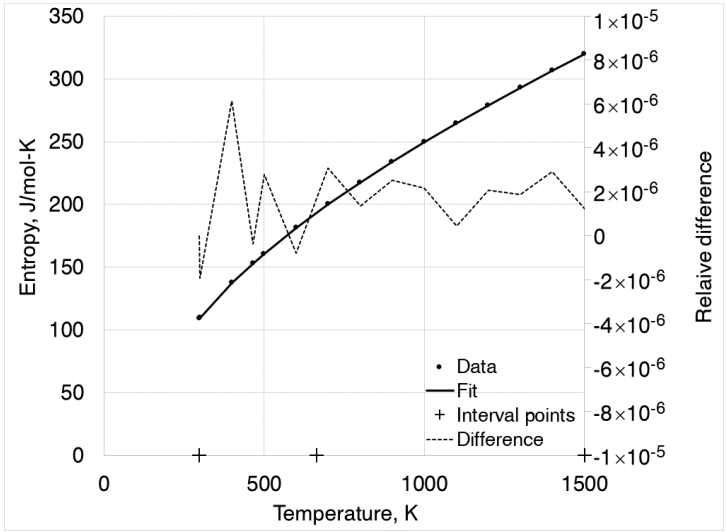


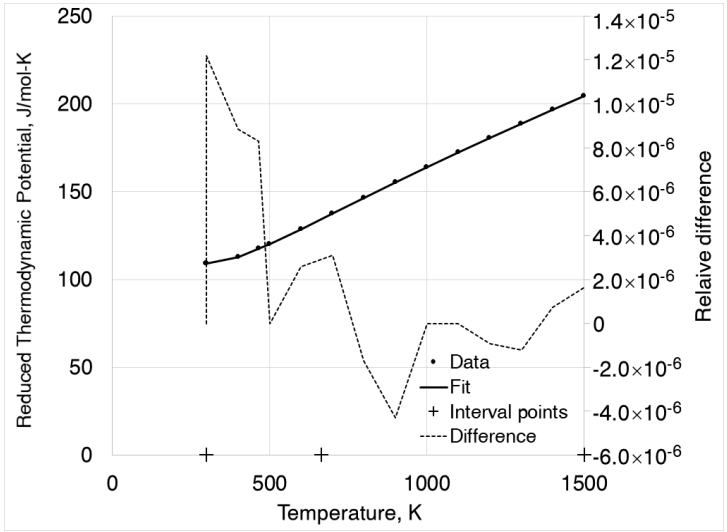
Aluminum Chloride ($AlCl_3$) – Al1Cl3(cr)

MW = 133.341 g/mol – $\Delta_f H^o$(298.15 K) = -705.632 kJ/mol – S^o(298.15 K) = 109.286 J/mol-K – p^o = 0.1 MPa



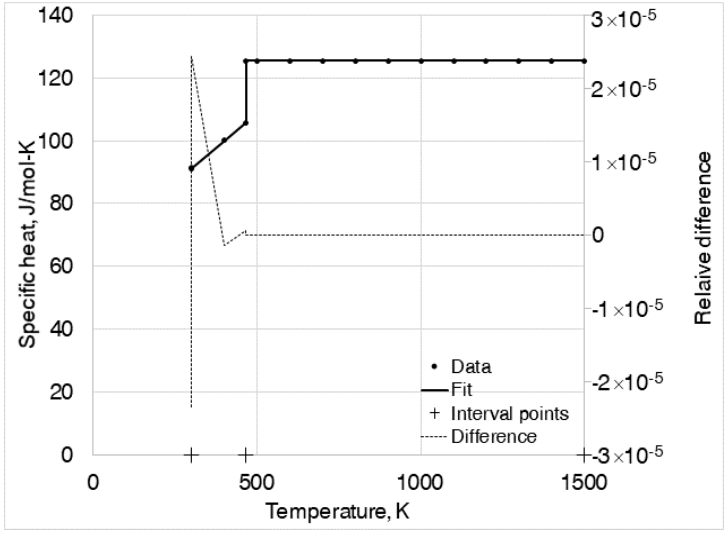


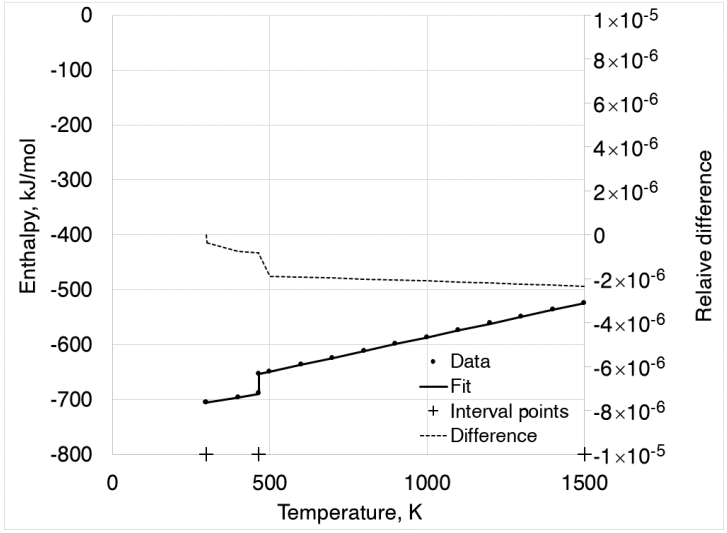


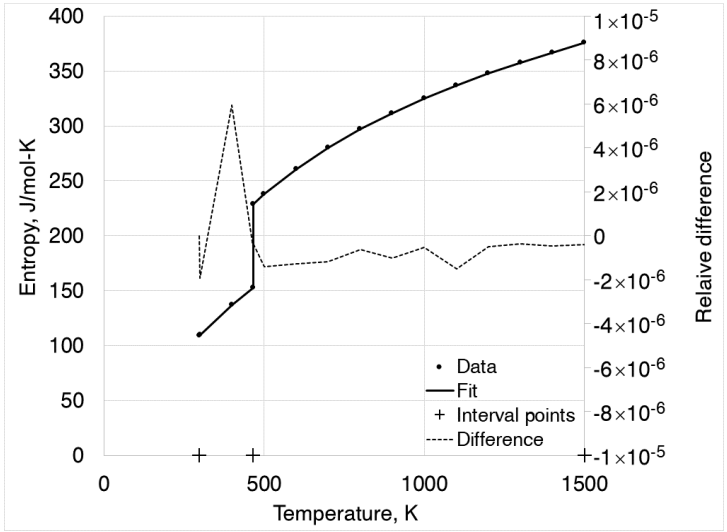


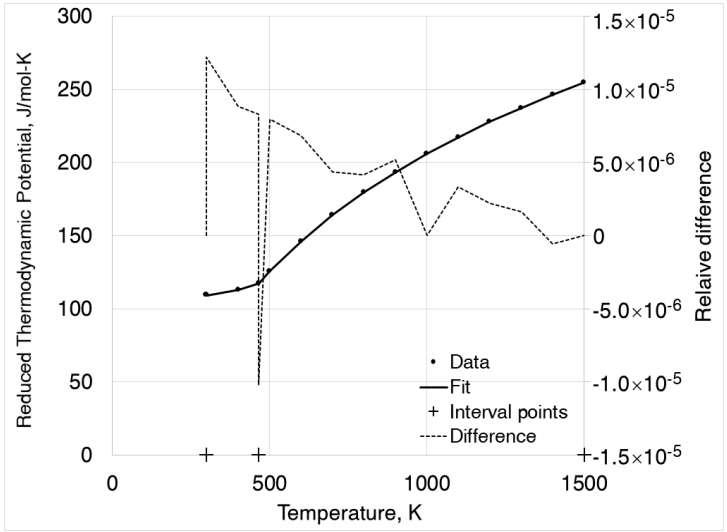
Aluminum Chloride ($AlCl_3$) – Al1Cl3(cr,l)

MW = 133.341 g/mol – $\Delta_f H^o$(298.15 K) = -705.632 kJ/mol – S^o(298.15 K) = 109.286 J/mol-K – p^o = 0.1 MPa



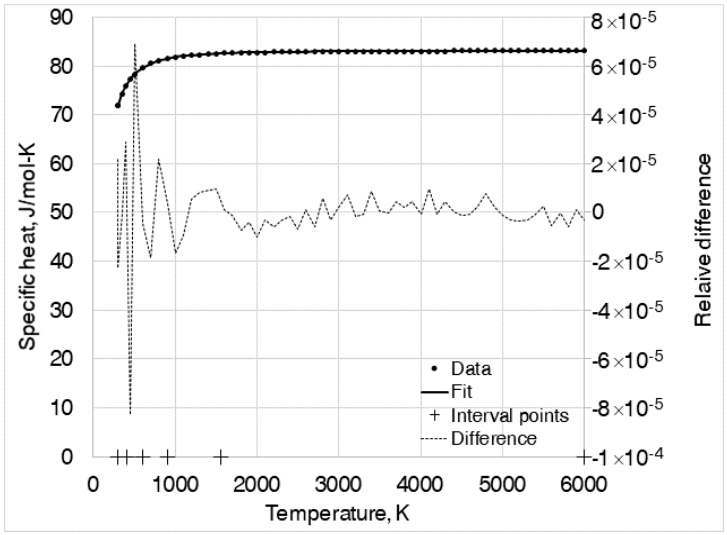


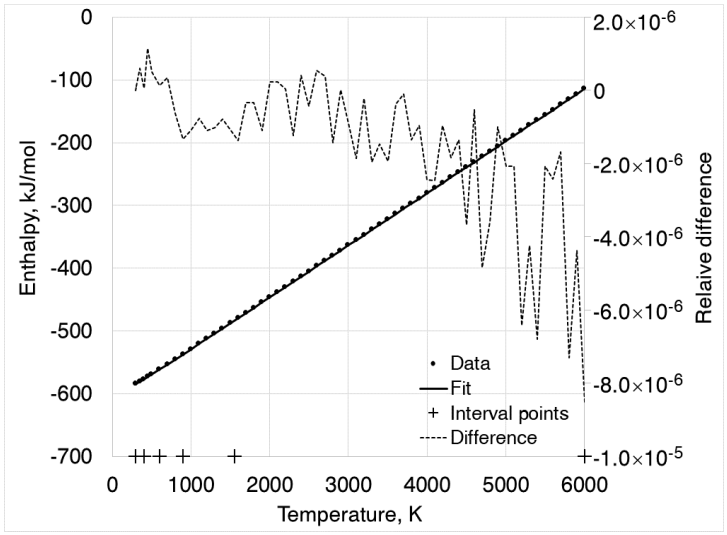


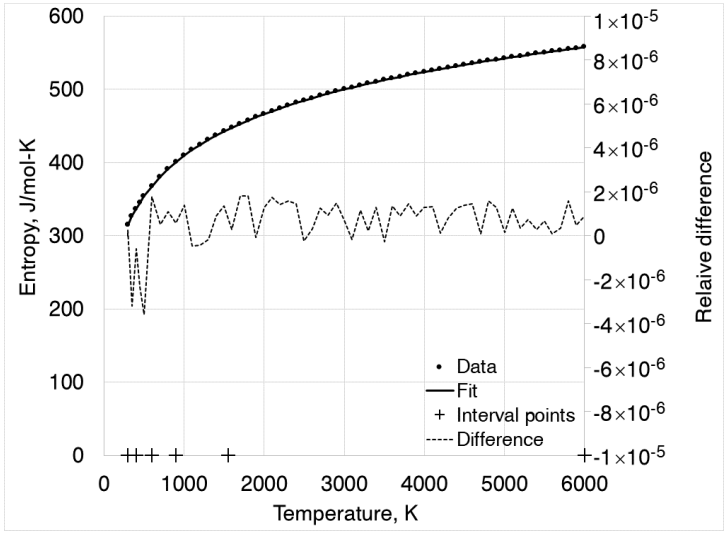


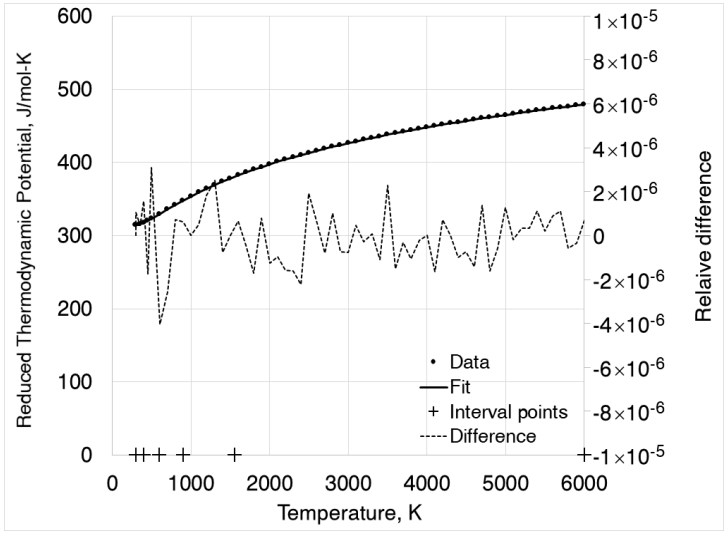
Aluminum Chloride ($AlCl_3$) – Al1Cl3(g)

MW = 133.341 g/mol – $\Delta_f H^o$(298.15 K) = -584.588 kJ/mol – S^o(298.15 K) = 314.494 J/mol-K – p^o = 0.1 MPa



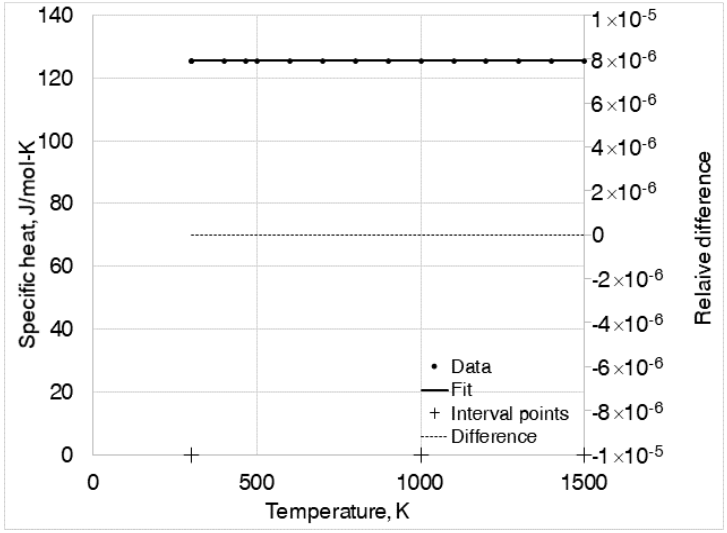


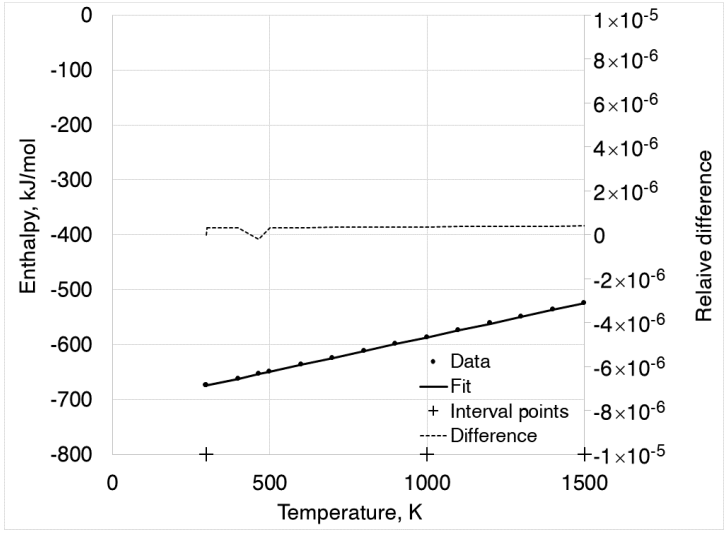


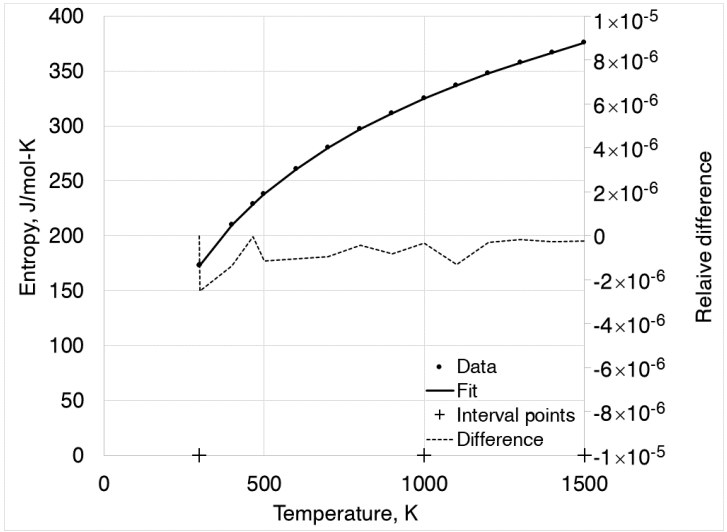


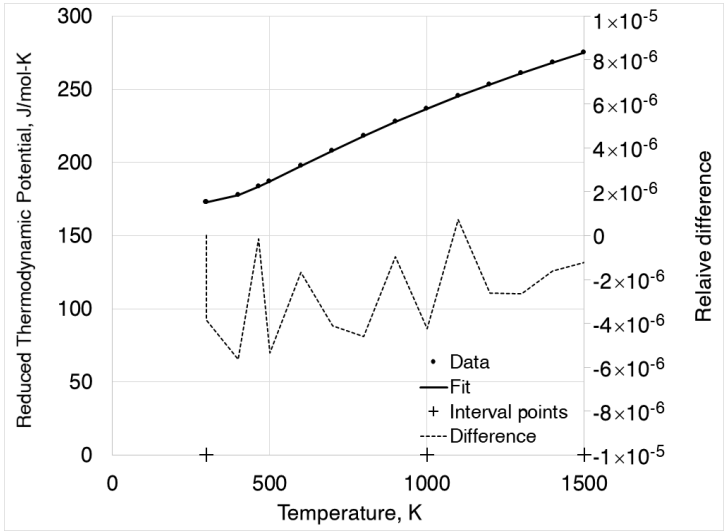
Aluminum Chloride ($AlCl_3$) – Al1Cl3(l)

MW = 133.341 g/mol – $\Delta_f H^o$(298.15 K) = -674.805 kJ/mol – S^o(298.15 K) = 172.908 J/mol-K – p^o = 0.1 MPa



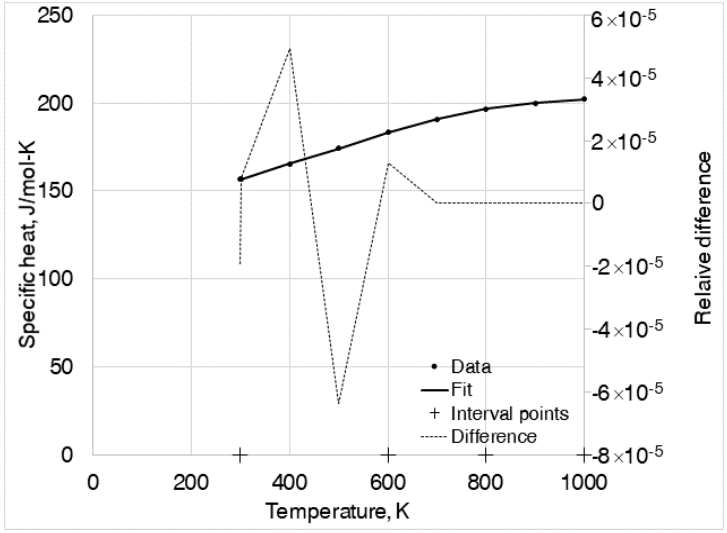


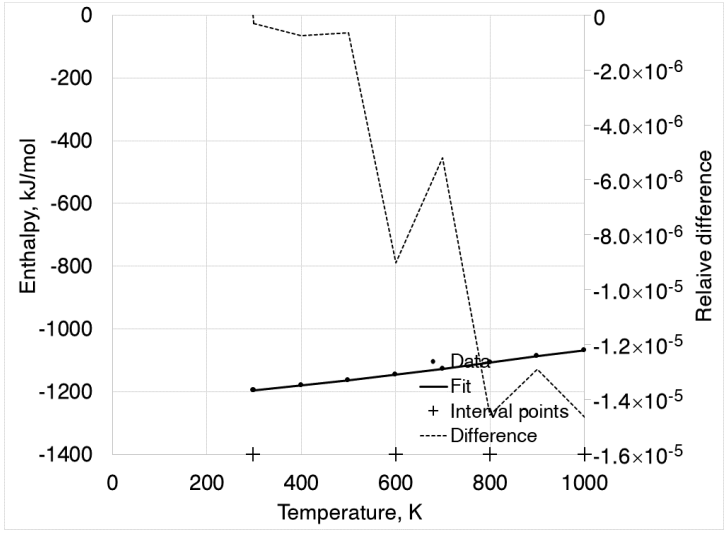


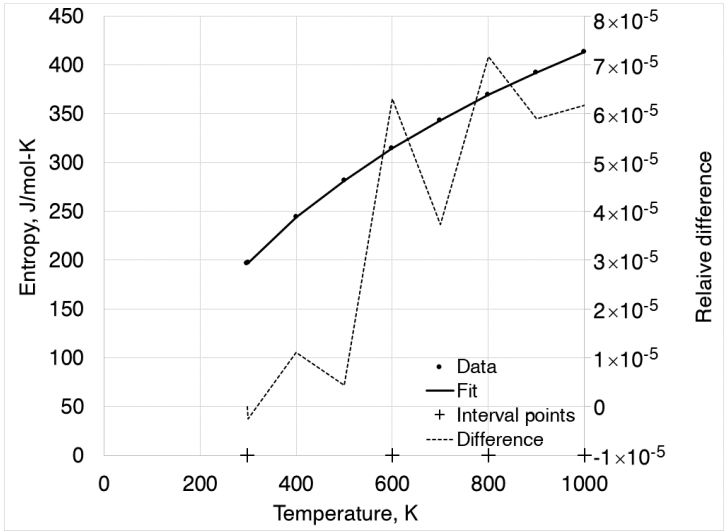


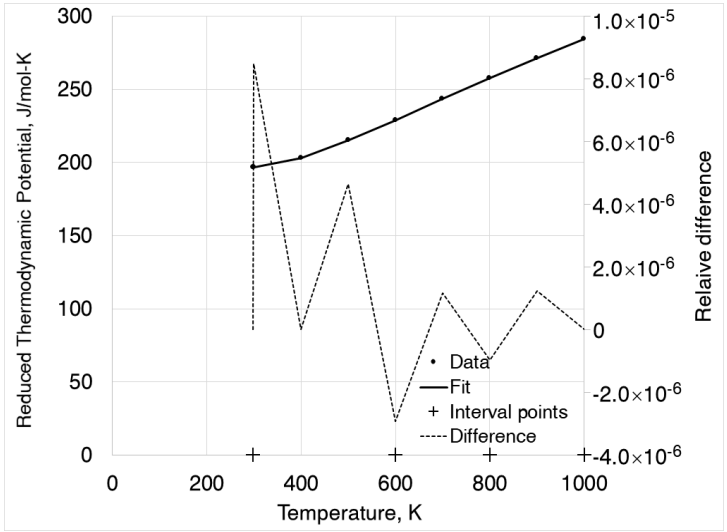
Potassium Tetrachloroaluminate ($KAlCl_4$) – Al1Cl4K1(cr)

MW = 207.892 g/mol – $\Delta_f H^o$(298.15 K) = -1196.624 kJ/mol – S^o(298.15 K) = 196.648 J/mol-K – p^o = 0.1 MPa



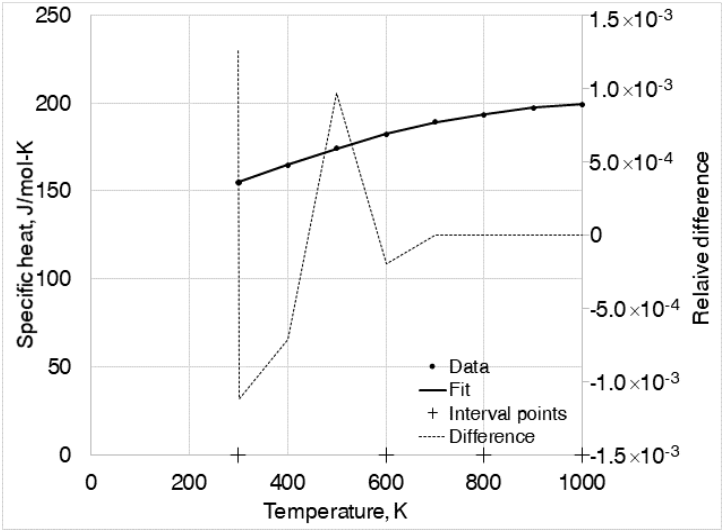


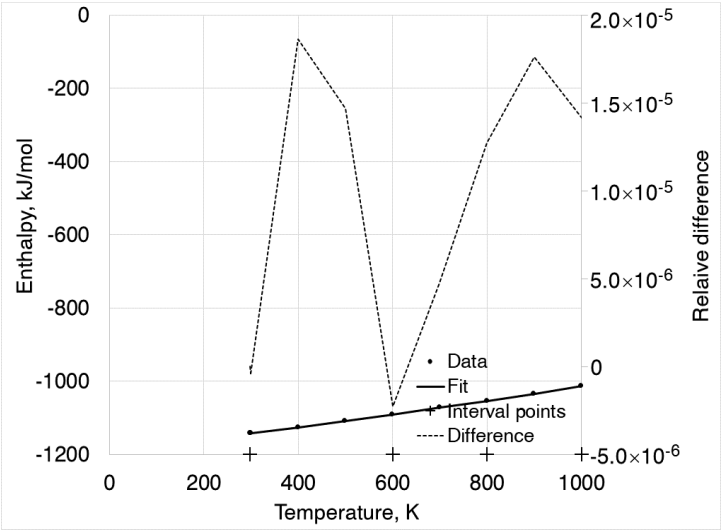


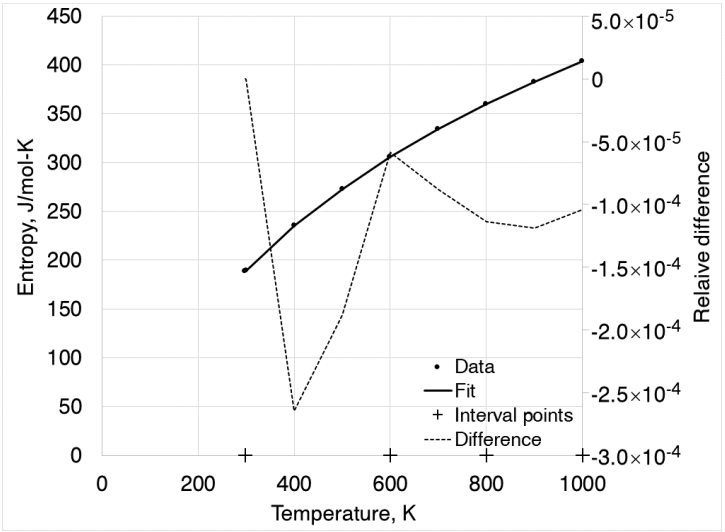


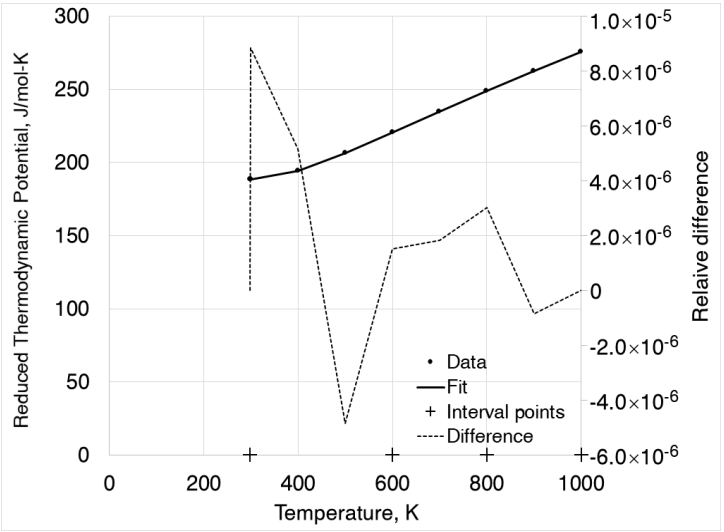
Sodium Tetrachloroaluminate ($NaAlCl_4$) – Al1Cl4Na1(cr)

MW = 191.783 g/mol – $\Delta_f H^o$(298.15 K) = -1142.232 kJ/mol – S^o(298.15 K) = 188.28 J/mol-K – p^o = 0.1 MPa



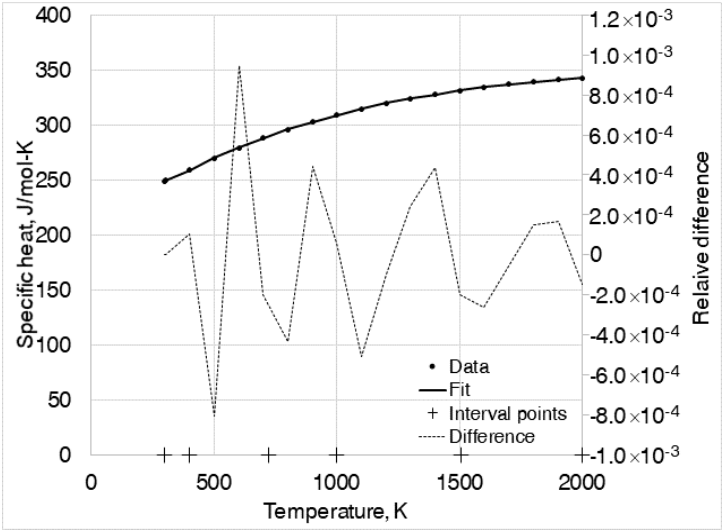


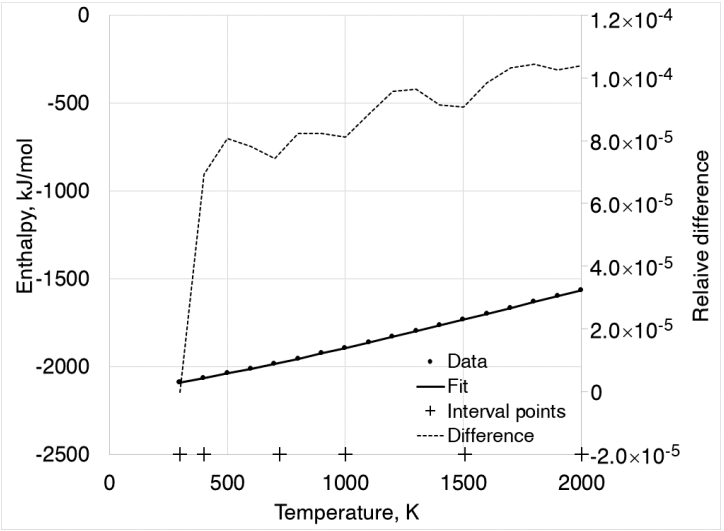


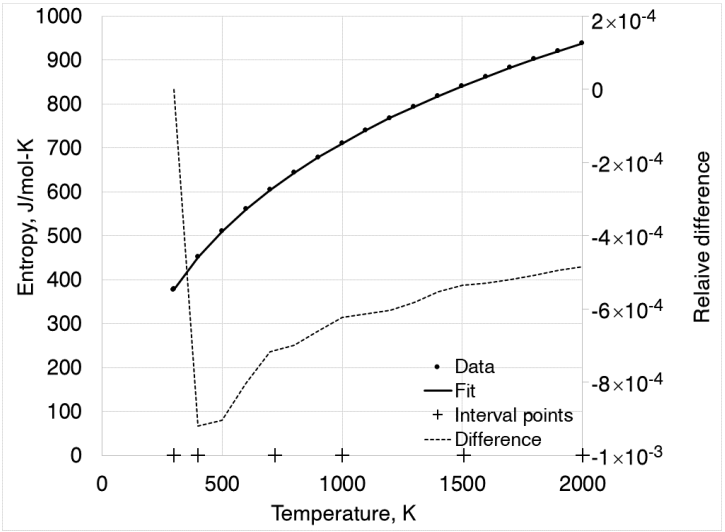


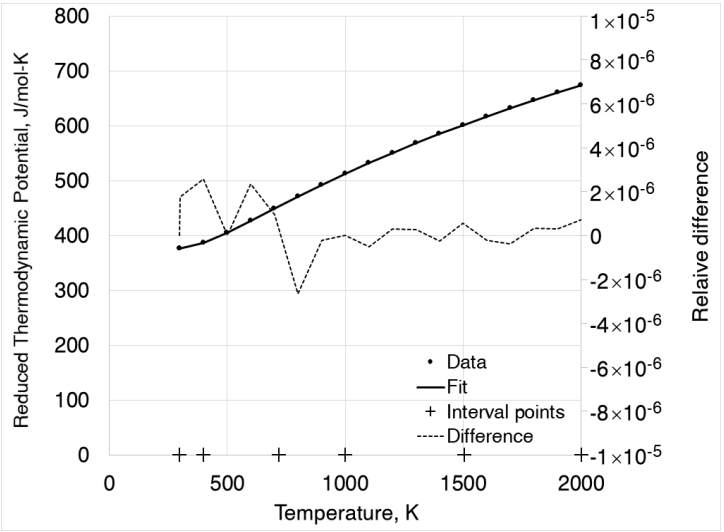
Potassium Hexachloroaluminate (K_3AlCl_6) – Al1Cl6K3(cr)

MW = 356.994 g/mol – $\Delta_f H^o$(298.15 K) = -2092 kJ/mol – S^o(298.15 K) = 376.56 J/mol-K – p^o = 0.1 MPa



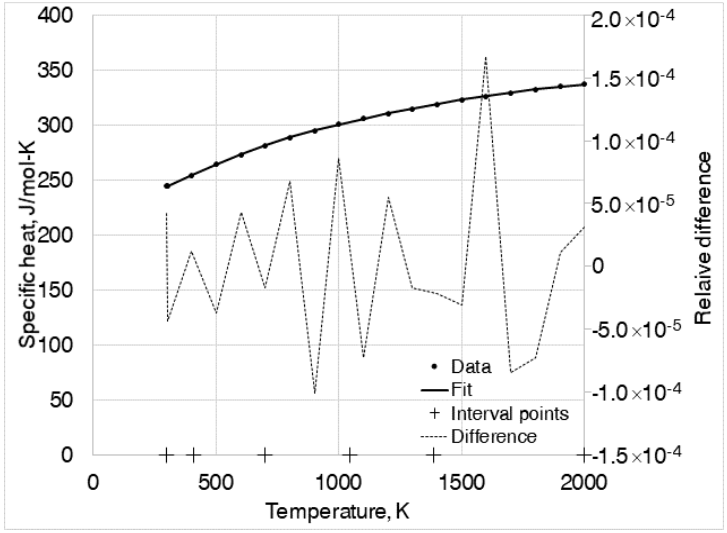


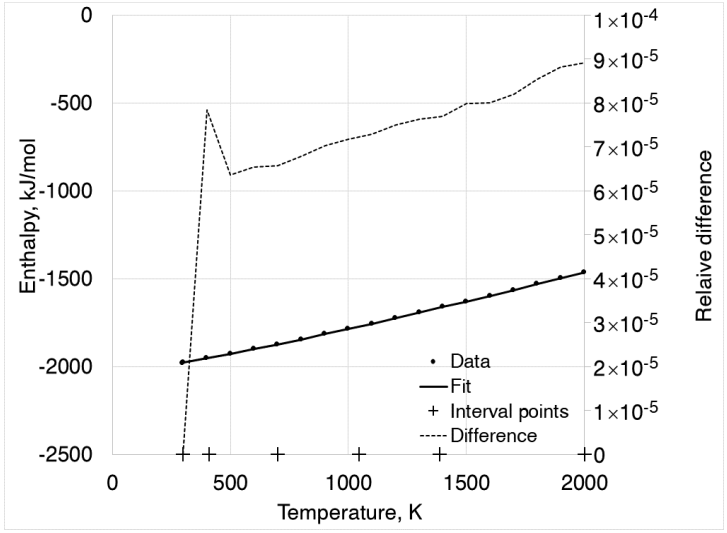


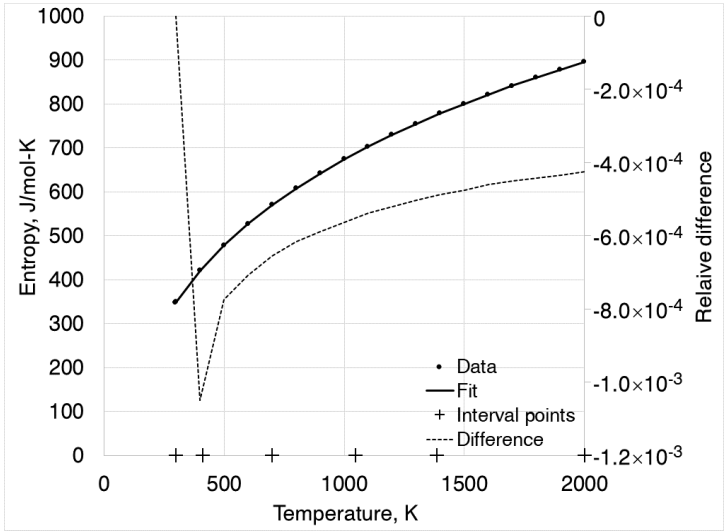


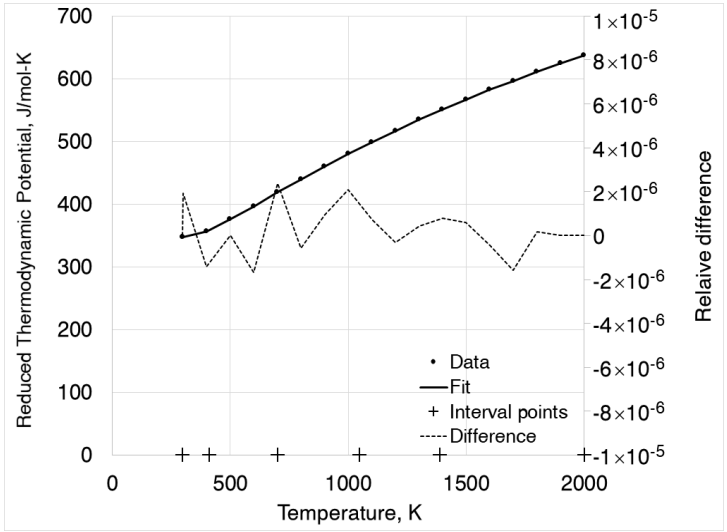
Sodium Hexachloroaluminate (Na_3AlCl_6) – Al1Cl6Na3(cr)

MW = 308.669 g/mol – $\Delta_f H^o$(298.15 K) = -1979.032 kJ/mol – S^o(298.15 K) = 347.272 J/mol-K – p^o = 0.1 MPa



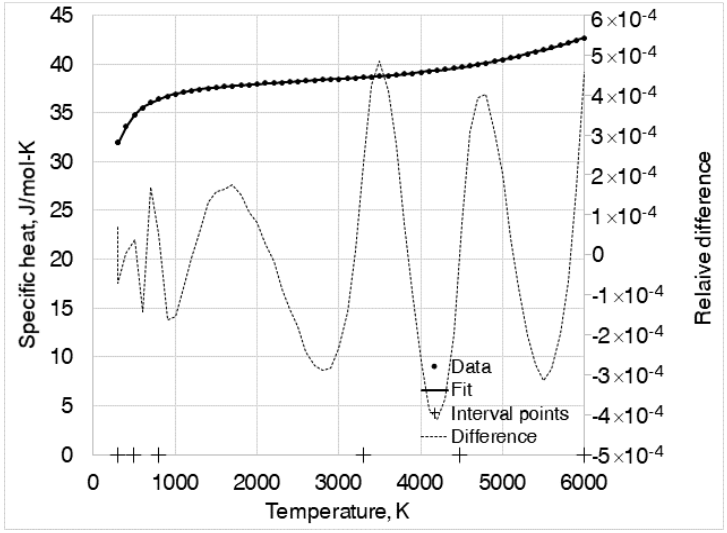


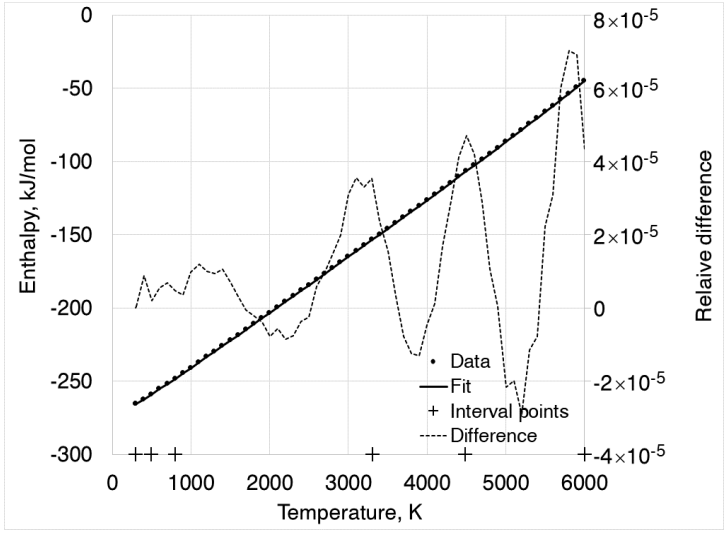


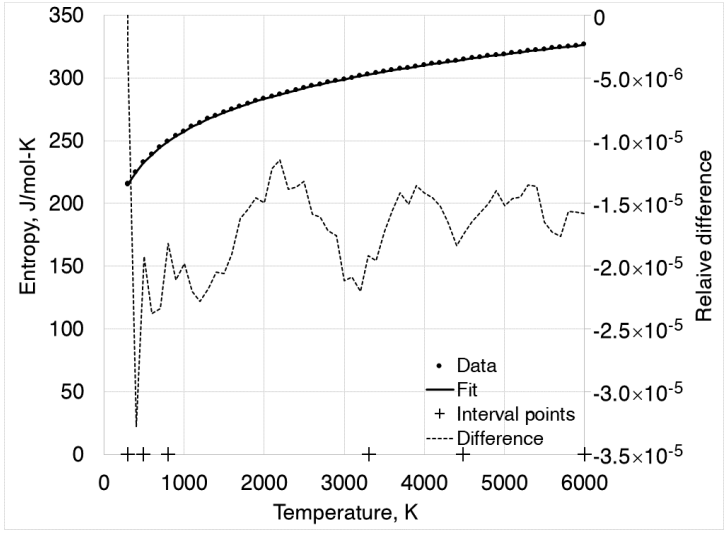


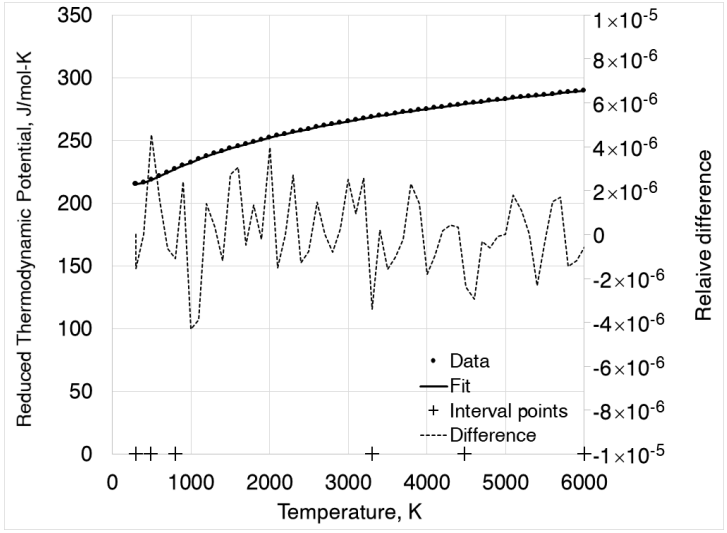
Aluminum Fluoride (AlF) – Al1F1(g)

MW = 45.9799 g/mol – $\Delta_f H^o$(298.15 K) = -265.684 kJ/mol – S^o(298.15 K) = 215.157 J/mol-K – p^o = 0.1 MPa



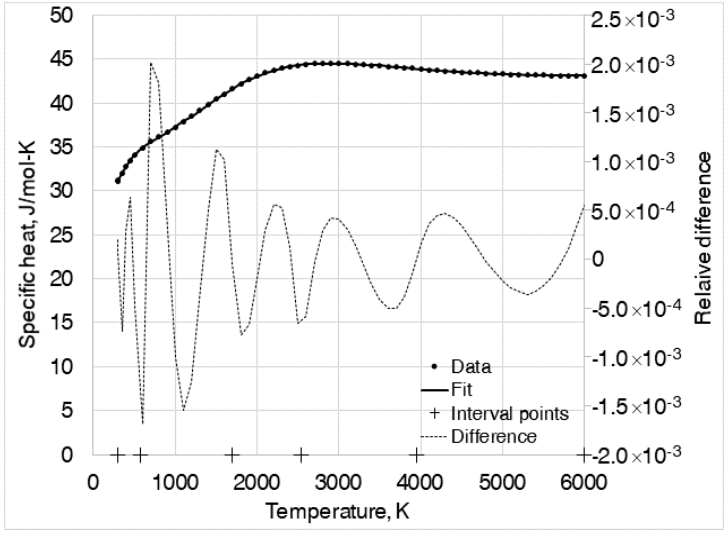


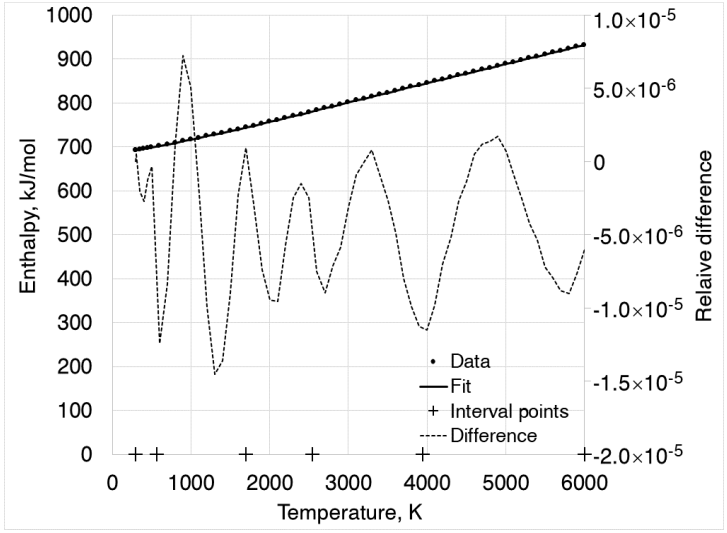


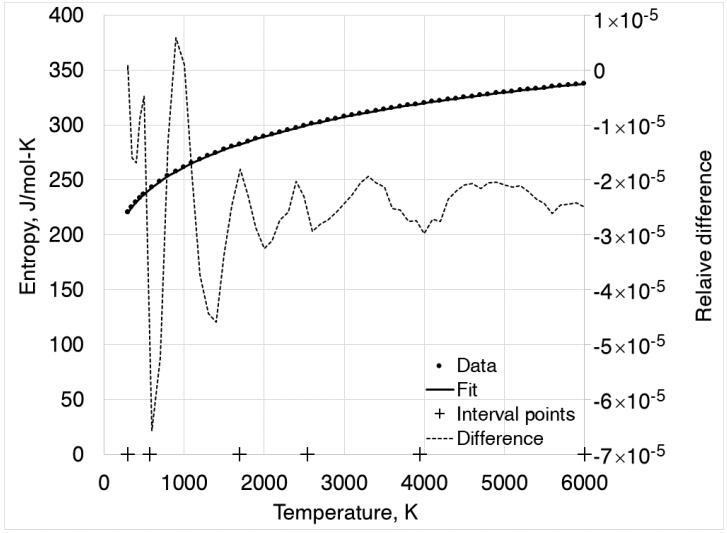


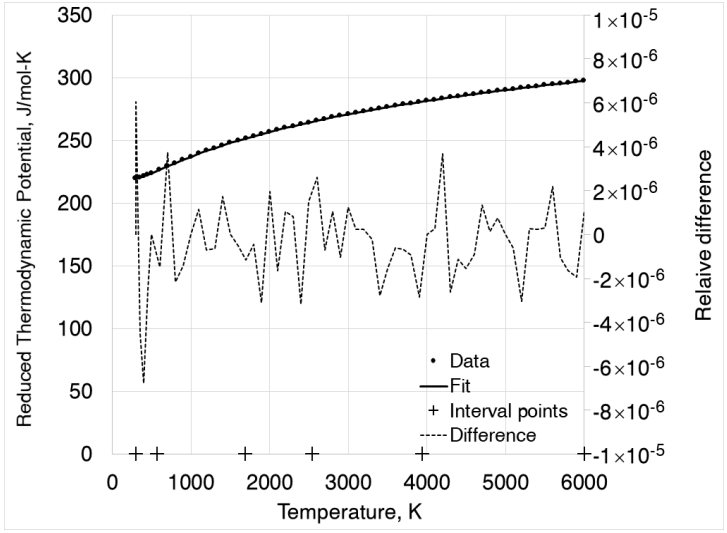
Aluminum Fluoride, Ion (AlF^+) – Al1F1+(g)

MW = 45.9794 g/mol – $\Delta_f H^o$(298.15 K) = 692.034 kJ/mol – S^o(298.15 K) = 220.065 J/mol-K – p^o = 0.1 MPa



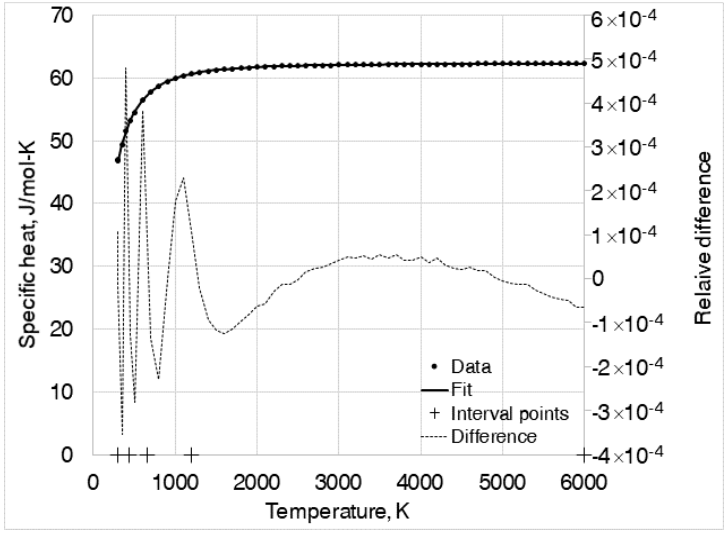


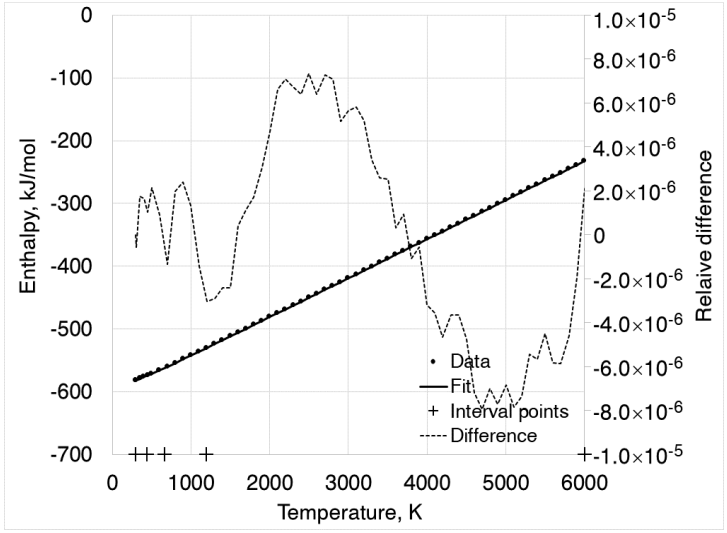


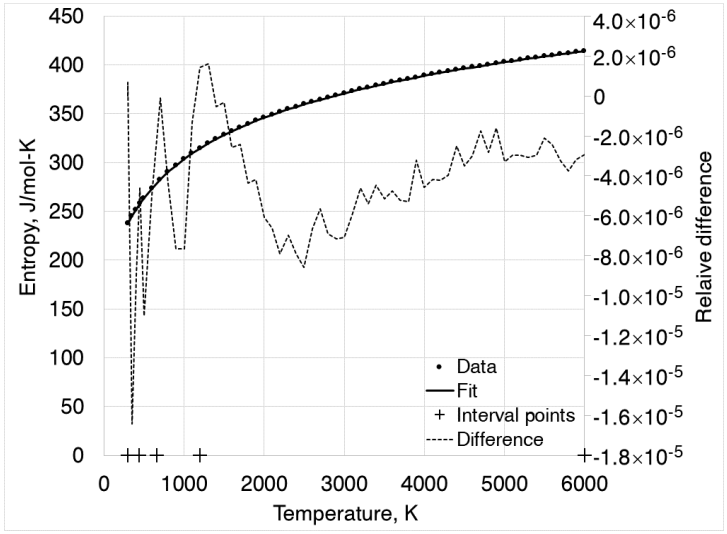


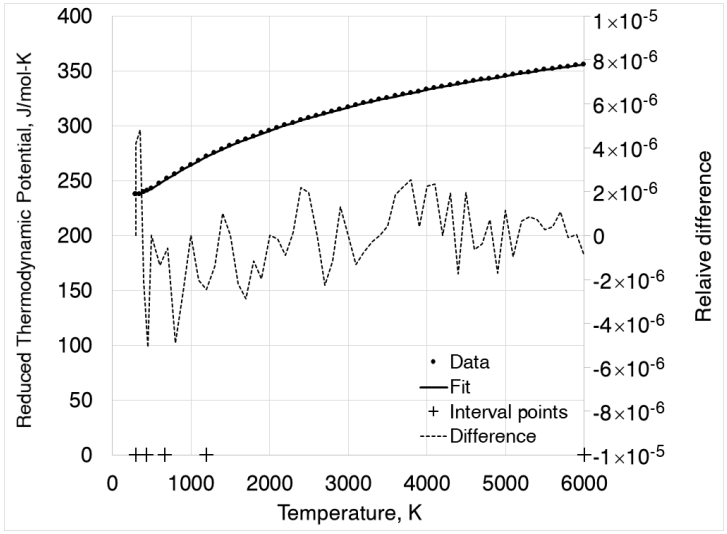
Aluminum Fluoride Oxide (OAlF) – Al1F1O1(g)

MW = 61.9793 g/mol – $\Delta_f H^o$(298.15 K) = -581.576 kJ/mol – S^o(298.15 K) = 237.316 J/mol-K – p^o = 0.1 MPa



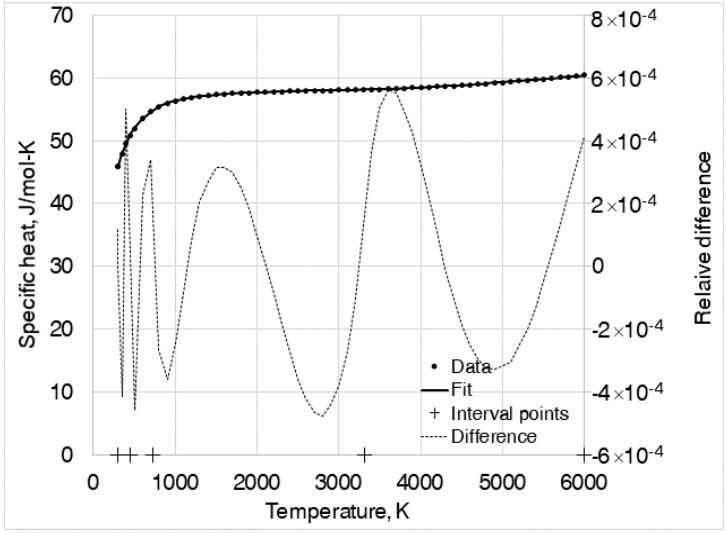


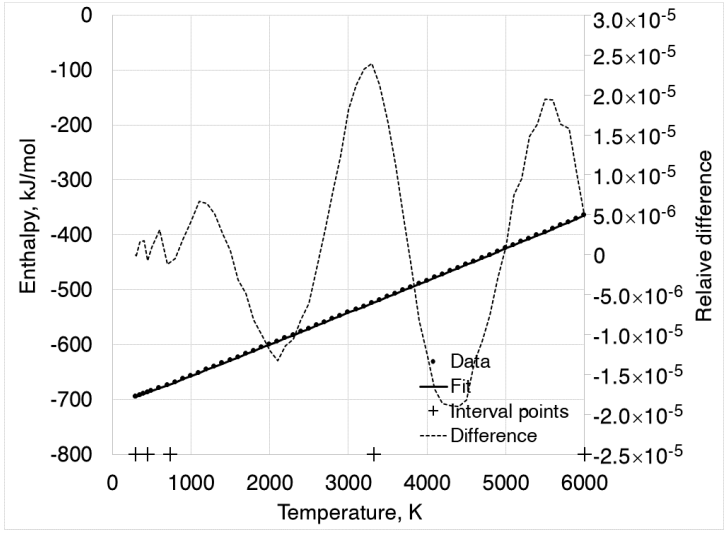


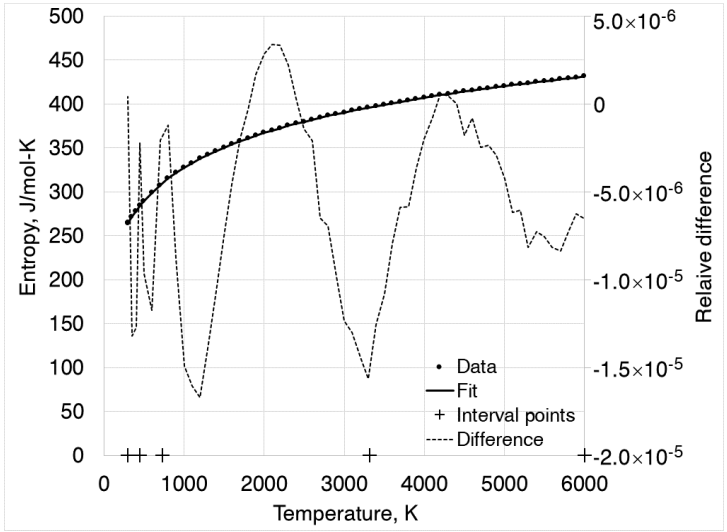


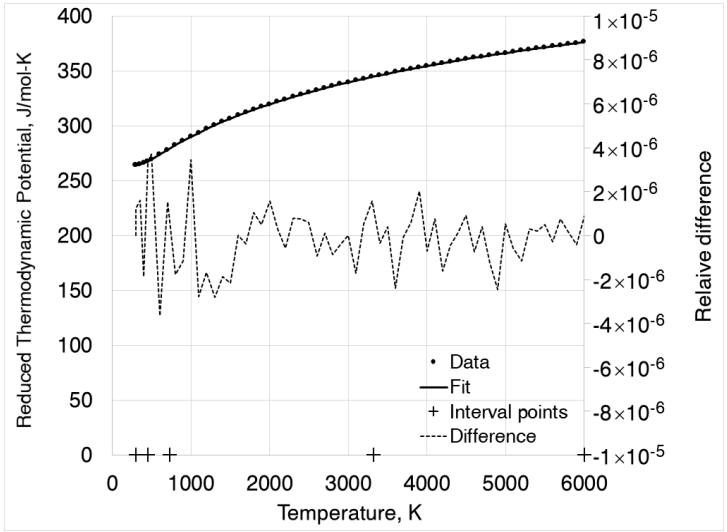
Aluminum Fluoride (AlF_2) – Al1F2(g)

MW = 64.9783 g/mol – $\Delta_f H^o$(298.15 K) = -694.544 kJ/mol – S^o(298.15 K) = 264.209 J/mol-K – p^o = 0.1 MPa



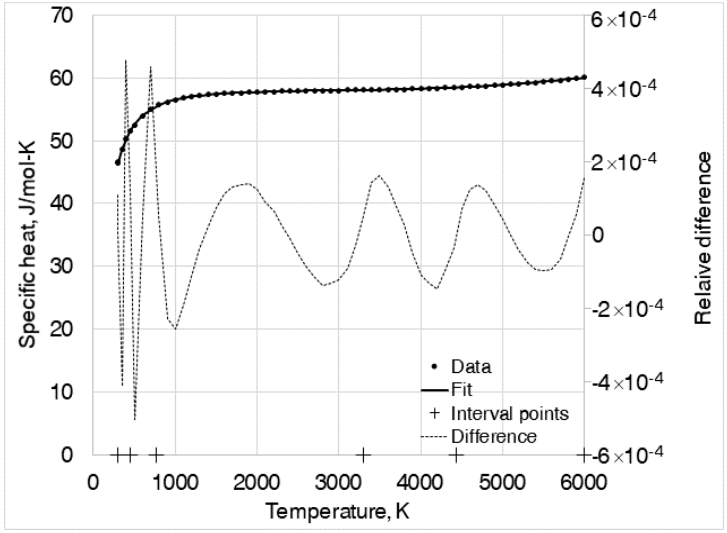


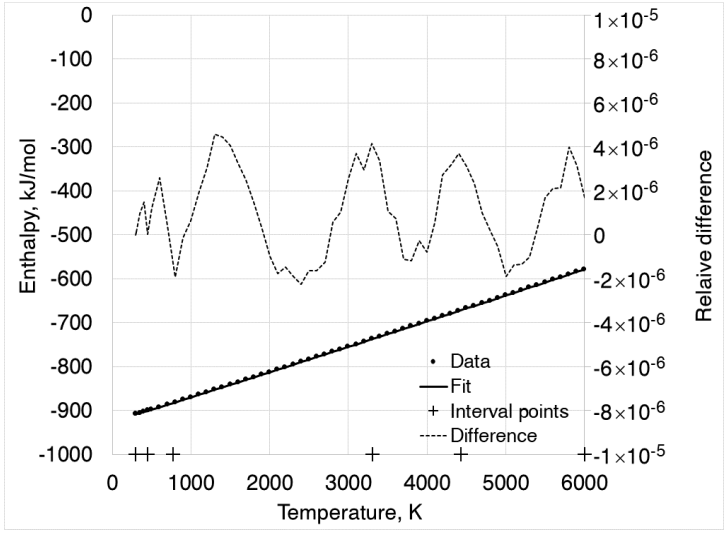


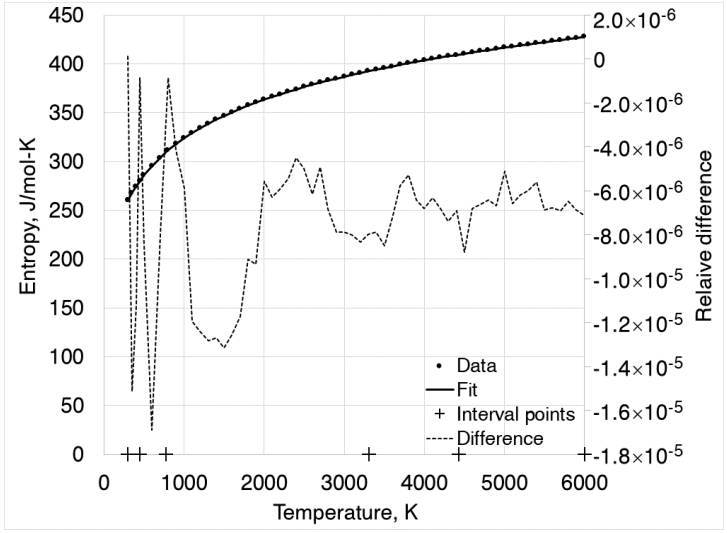


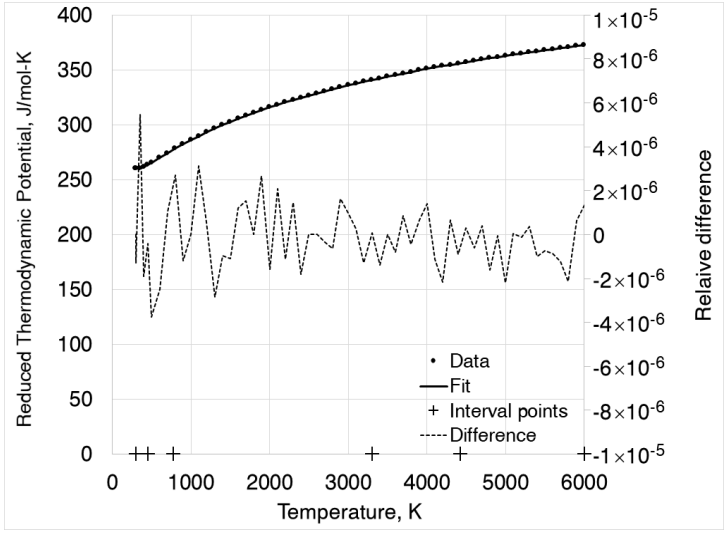
Aluminum Fluoride, Ion (AlF_2^-) – Al1F2-(g)

MW = 64.9789 g/mol – $\Delta_f H^o$(298.15 K) = -907.928 kJ/mol – S^o(298.15 K) = 260.046 J/mol-K – p^o = 0.1 MPa



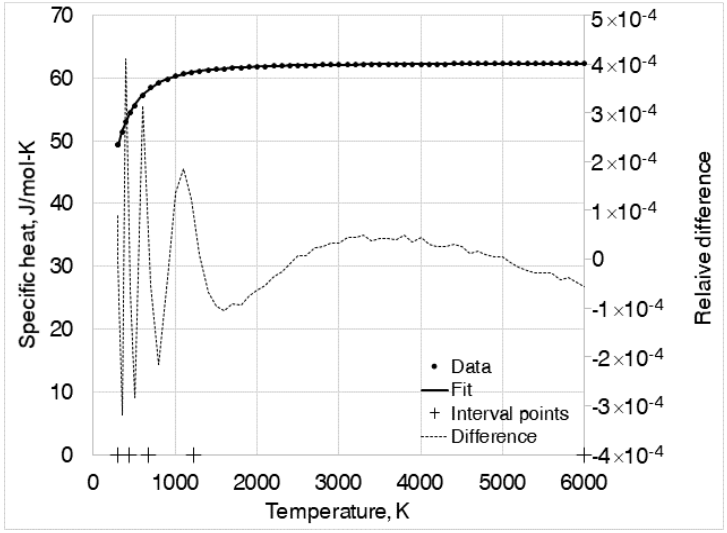


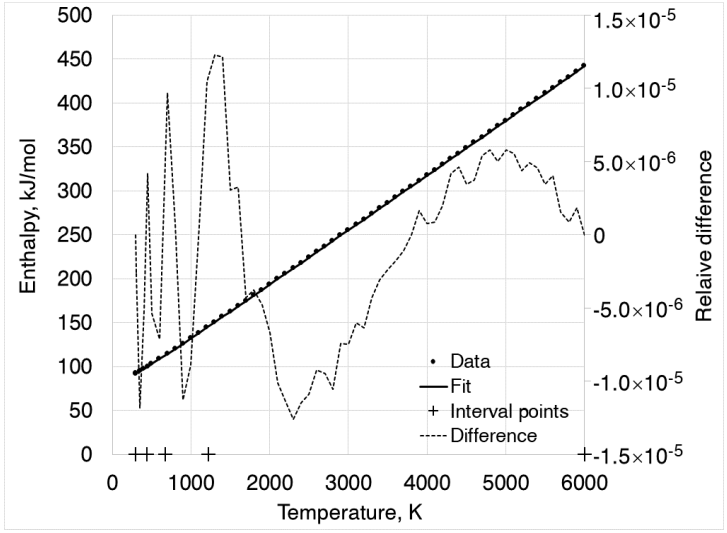


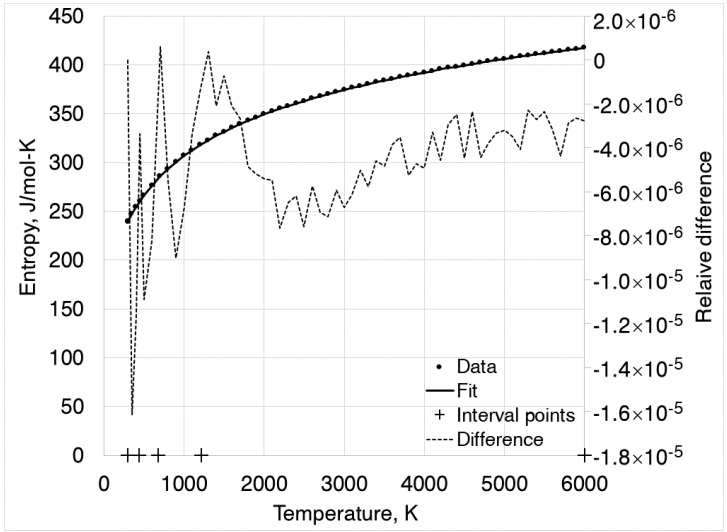


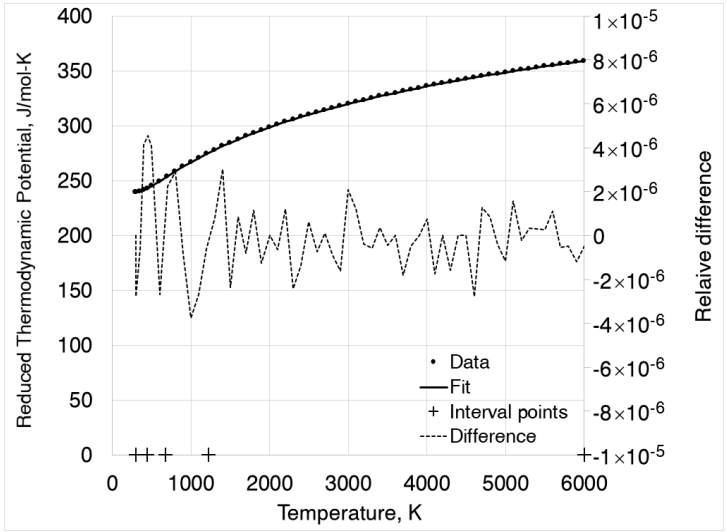
Aluminum Fluoride, Ion (AlF_2^+) – Al1F2+(g)

MW = 64.9778 g/mol – $\Delta_f H^o$(298.15 K) = 92.048 kJ/mol – S^o(298.15 K) = 239.299 J/mol-K – p^o = 0.1 MPa



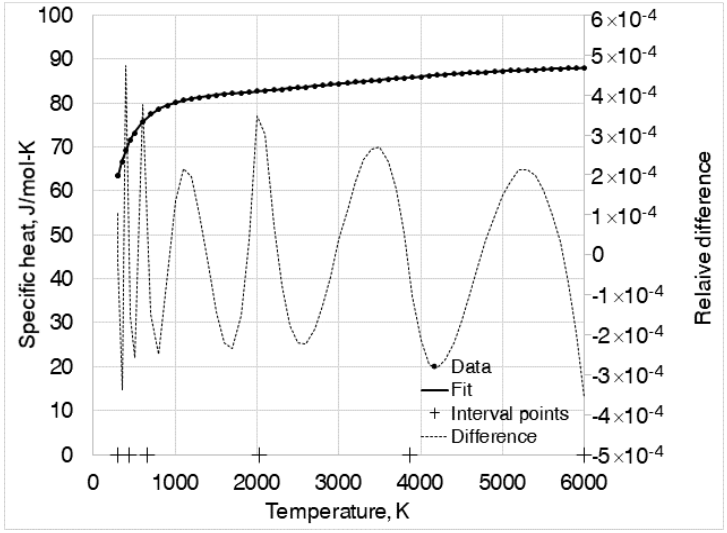


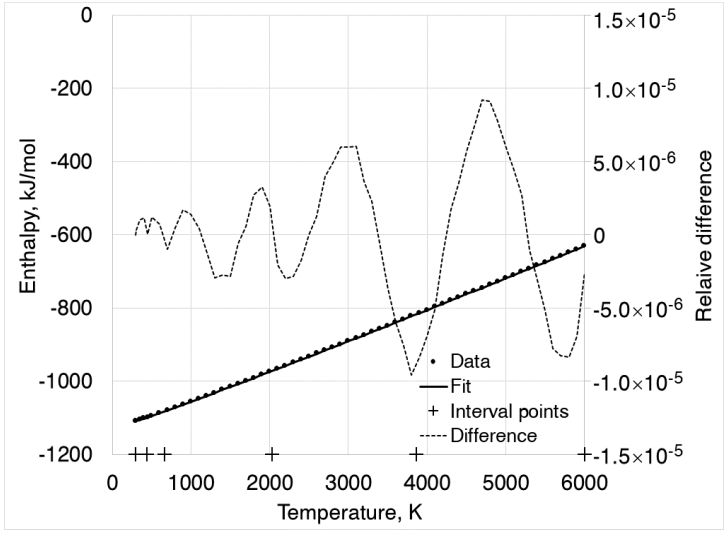


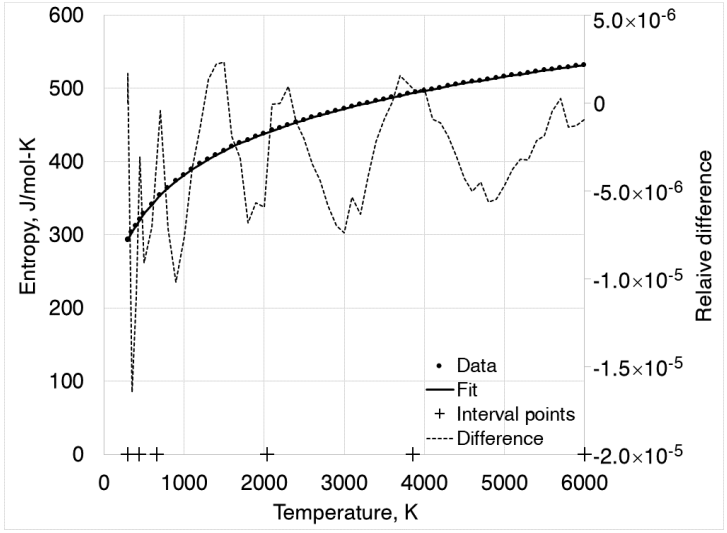


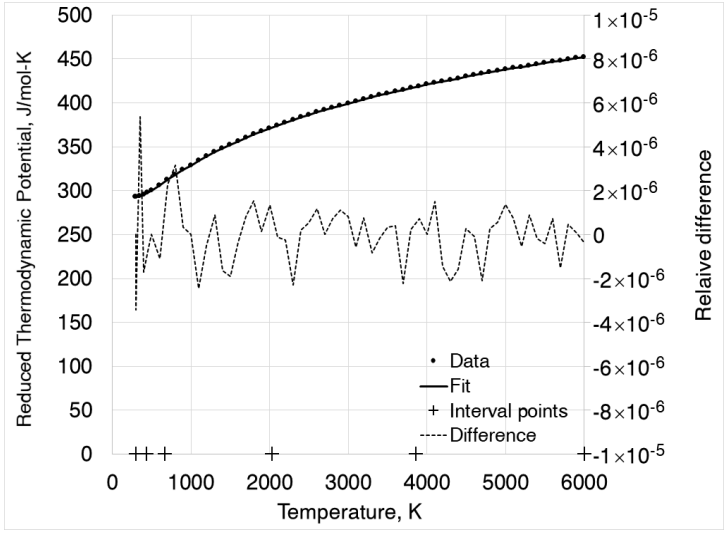
Aluminum Fluoride Oxide ($OAlF_2$) – Al1F2O1(g)

MW = 80.9777 g/mol – $\Delta_f H^o$(298.15 K) = -1108.76 kJ/mol – S^o(298.15 K) = 292.705 J/mol-K – p^o = 0.1 MPa



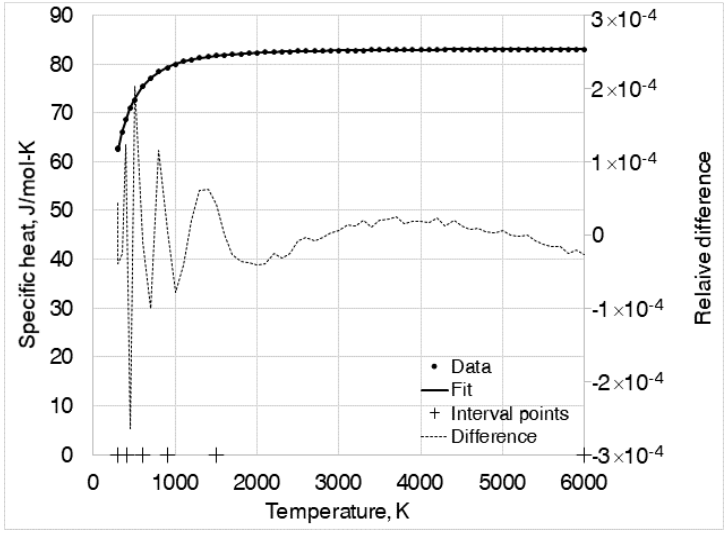


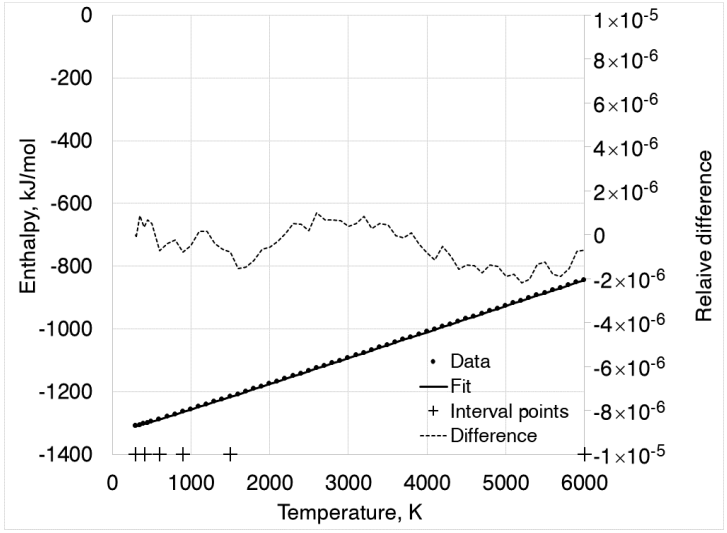


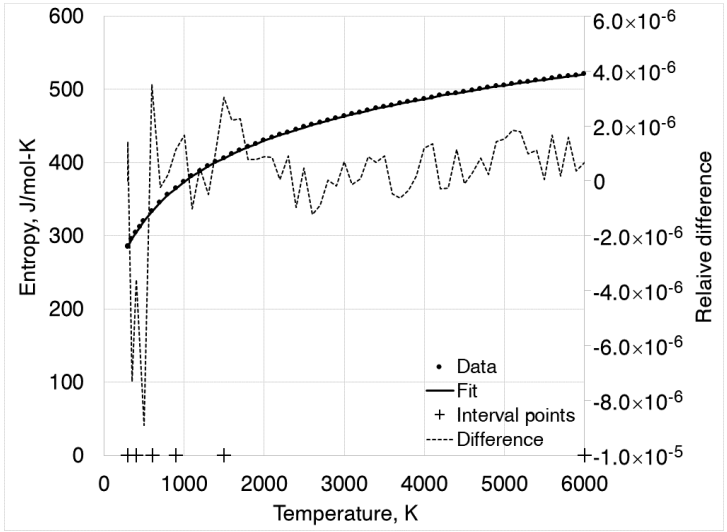


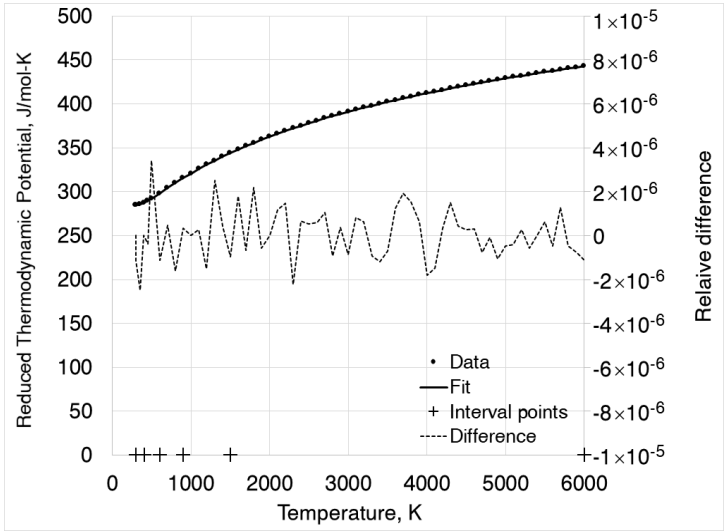
Aluminum Fluoride Oxide, Ion ($OAlF_2^-$) – Al1F2O1-(g)

MW = 80.9783 g/mol – $\Delta_f H^o$(298.15 K) = -1309.592 kJ/mol – S^o(298.15 K) = 284.87 J/mol-K – p^o = 0.1 MPa



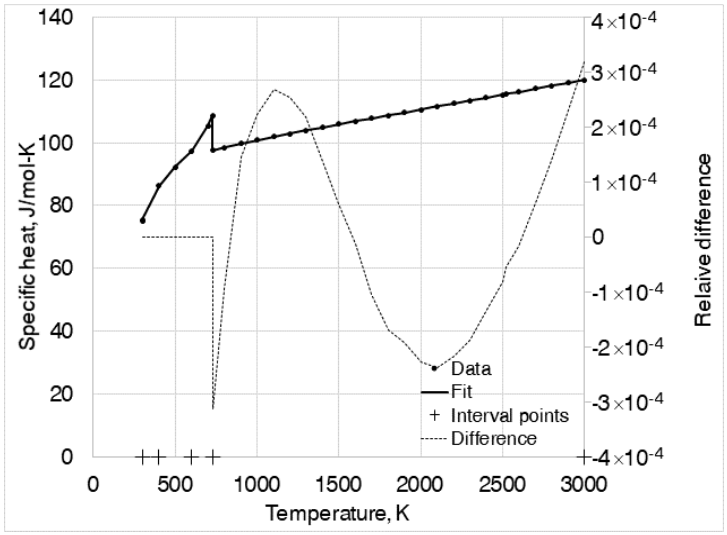


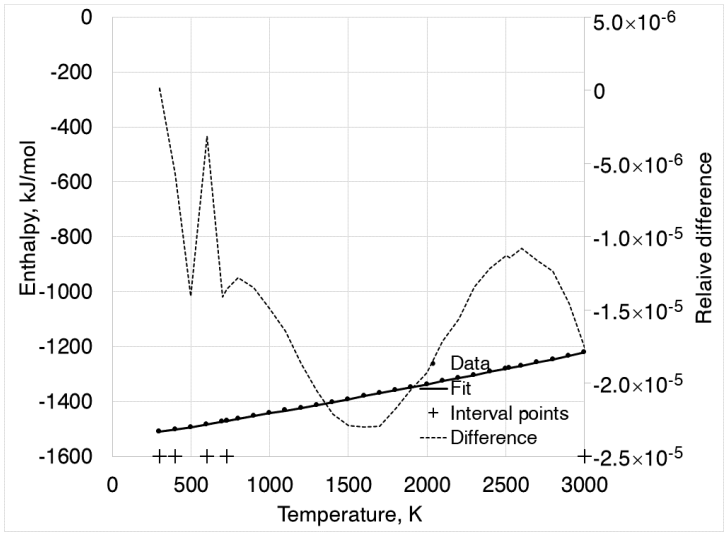


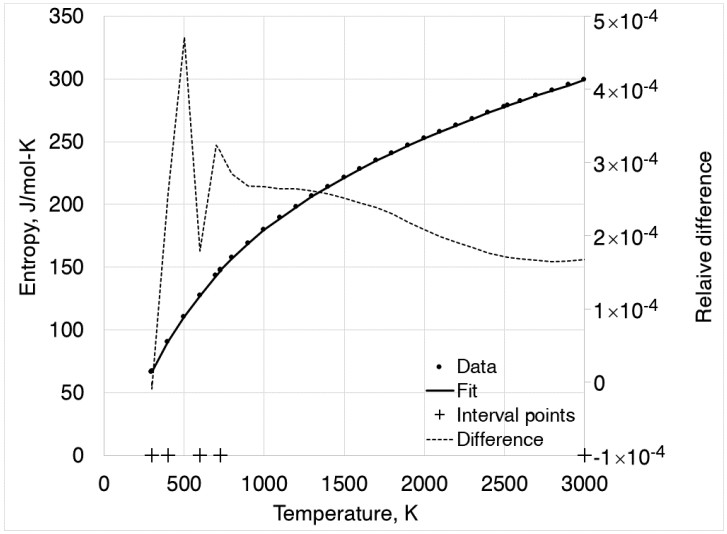


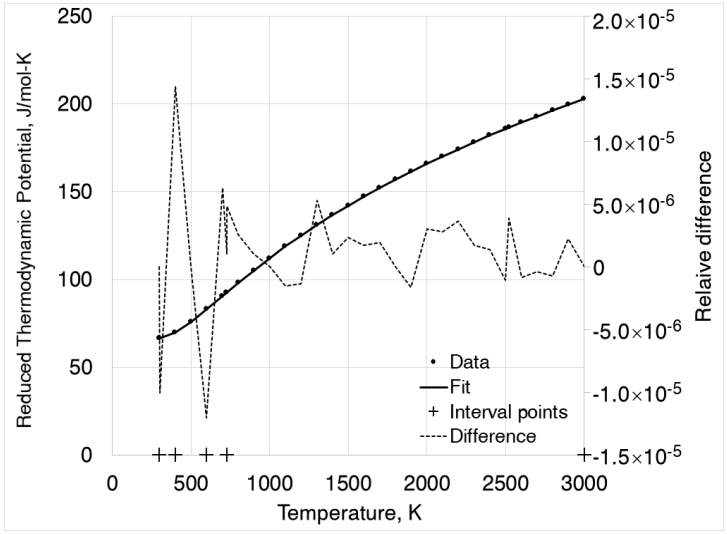
Aluminum Fluoride (AlF_3) – Al1F3(cr)

MW = 83.9767 g/mol – $\Delta_f H^o$(298.15 K) = -1510.424 kJ/mol – S^o(298.15 K) = 66.484 J/mol-K – p^o = 0.1 MPa



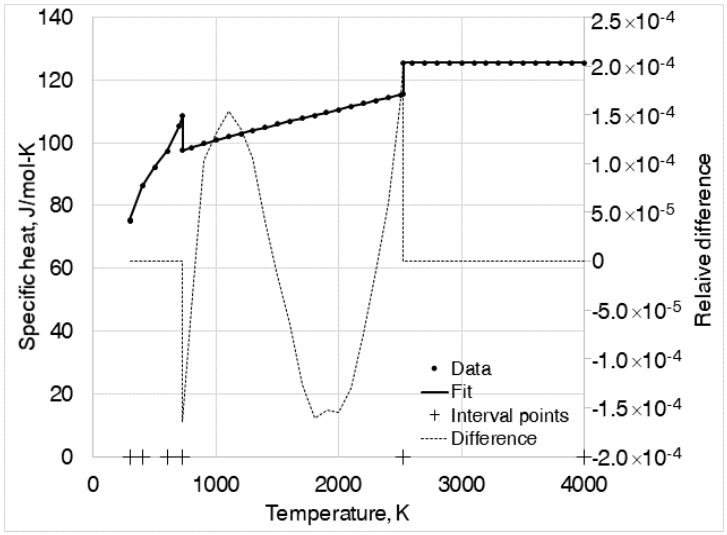


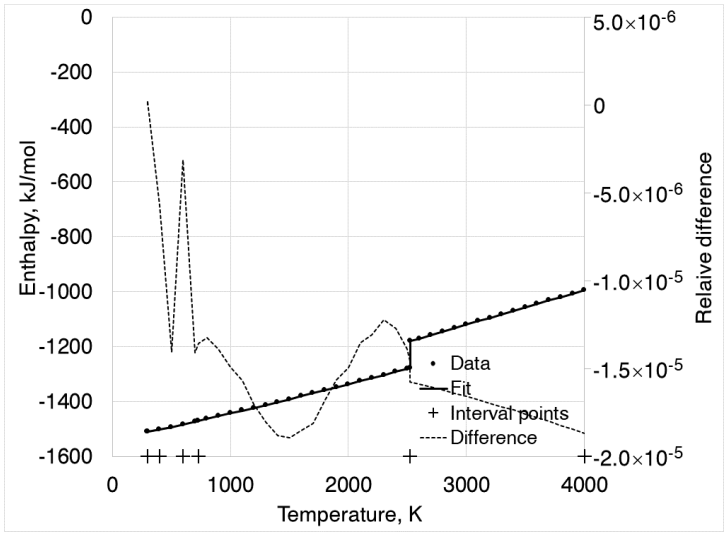


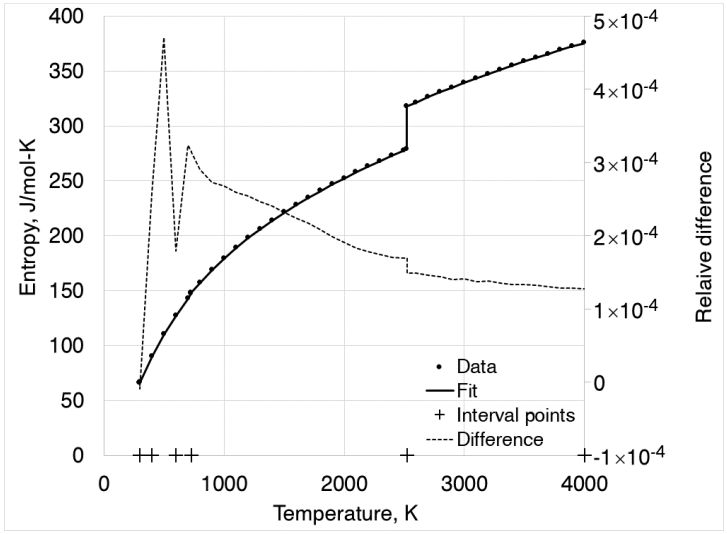


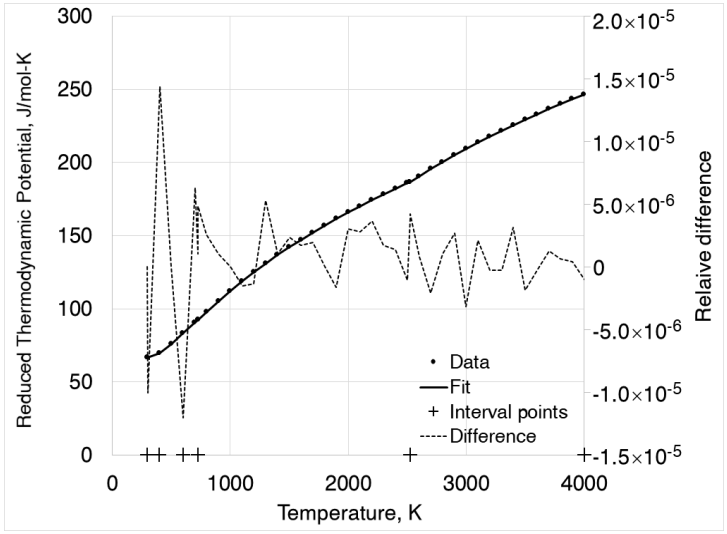
Aluminum Fluoride (AlF_3) – Al1F3(cr,l)

MW = 83.9767 g/mol – $\Delta_f H^o$(298.15 K) = -1510.424 kJ/mol – S^o(298.15 K) = 66.484 J/mol-K – p^o = 0.1 MPa



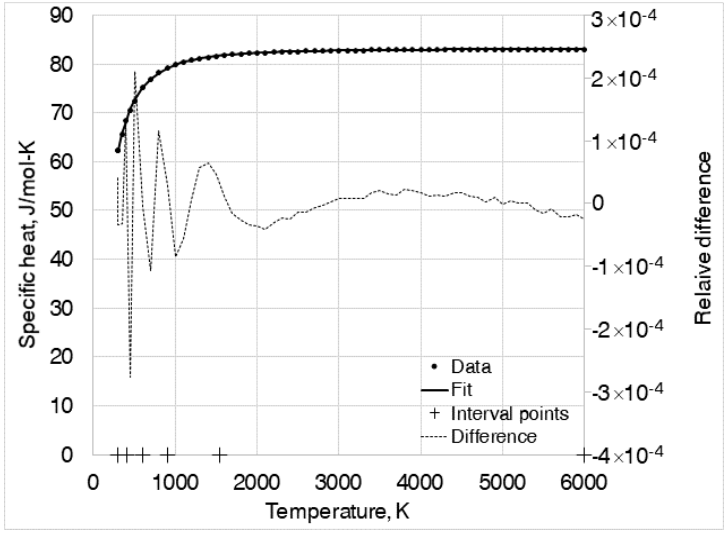


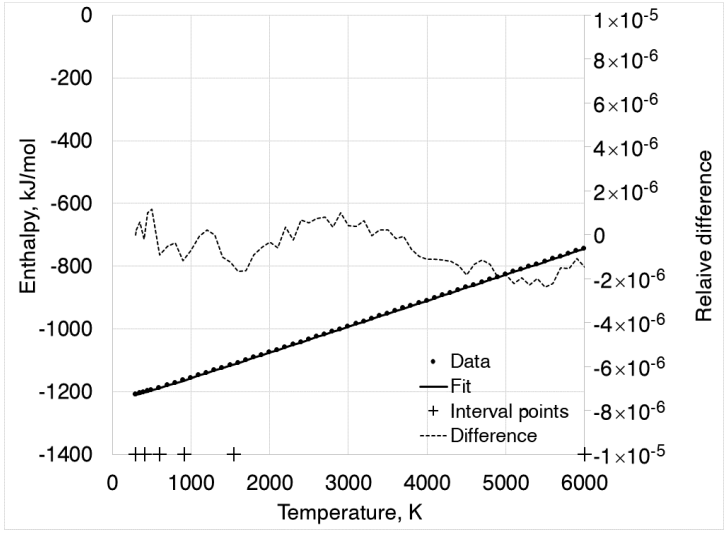


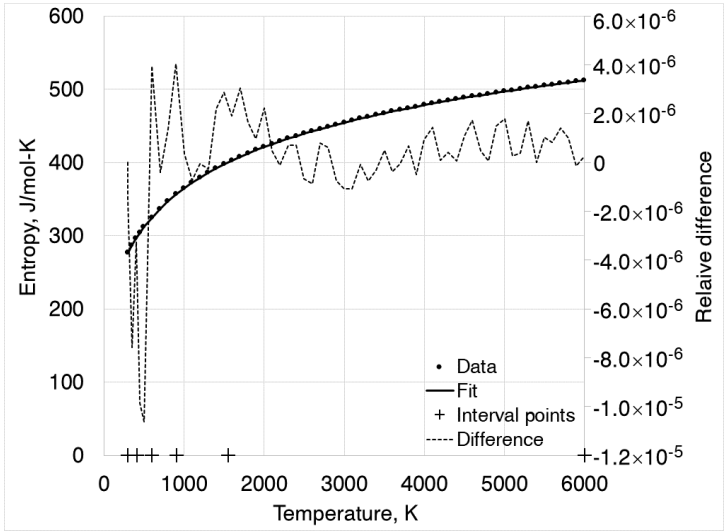


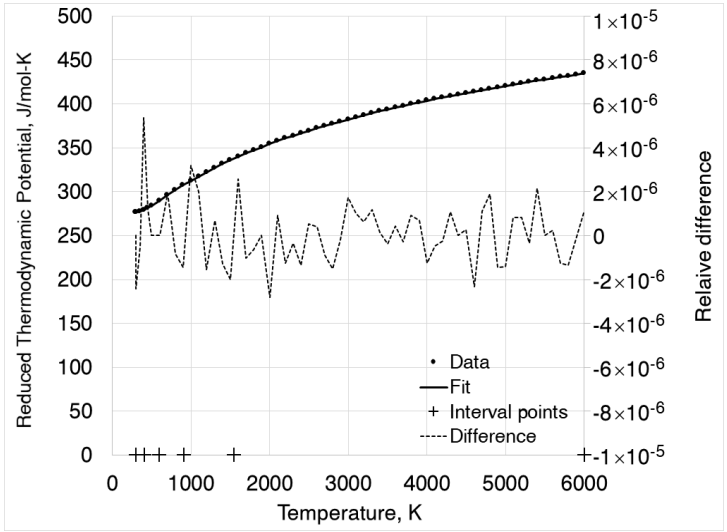
Aluminum Fluoride (AlF_3) – Al1F3(g)

MW = 83.9767 g/mol – $\Delta_f H^o$(298.15 K) = -1209.322 kJ/mol – S^o(298.15 K) = 276.691 J/mol-K – p^o = 0.1 MPa



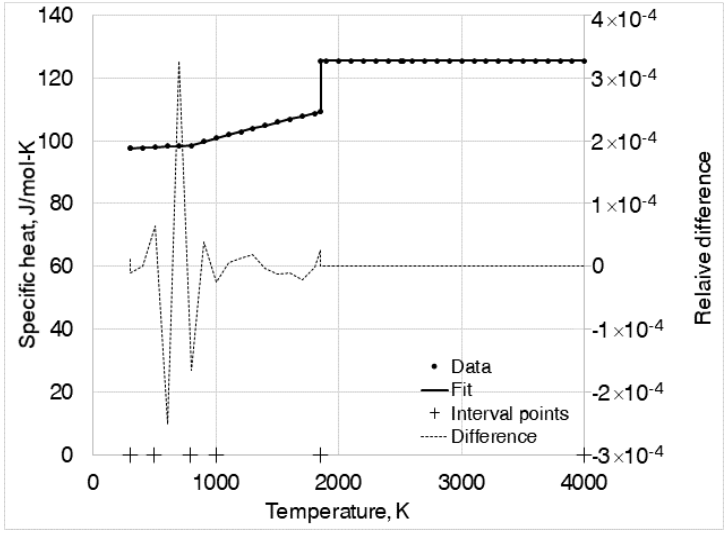


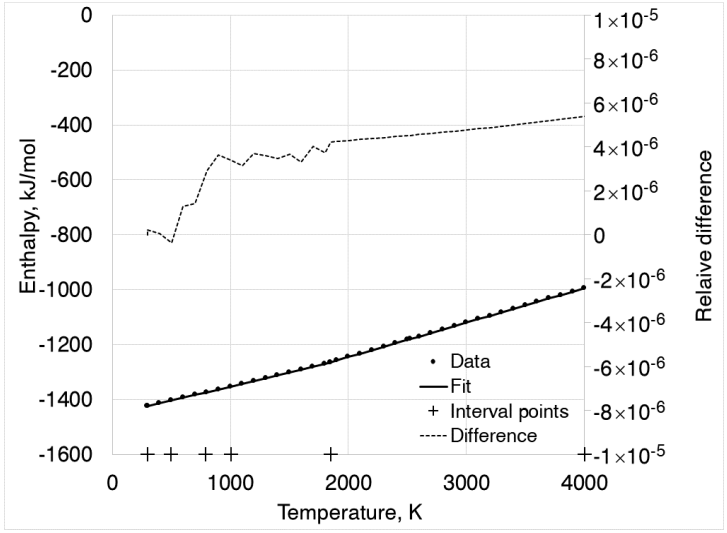


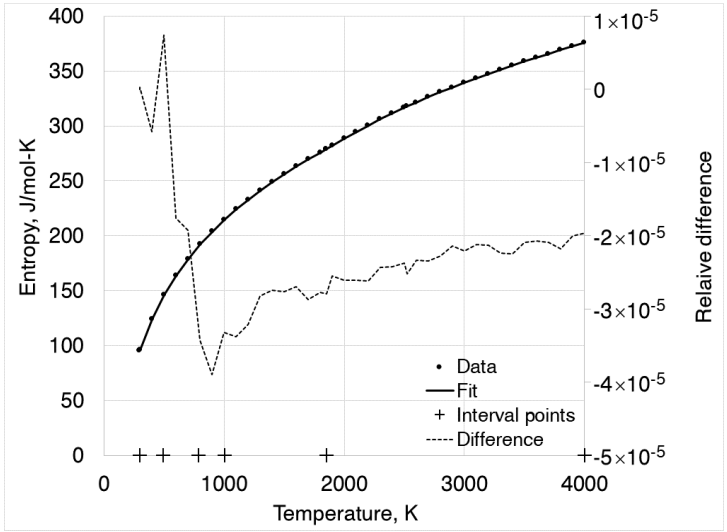


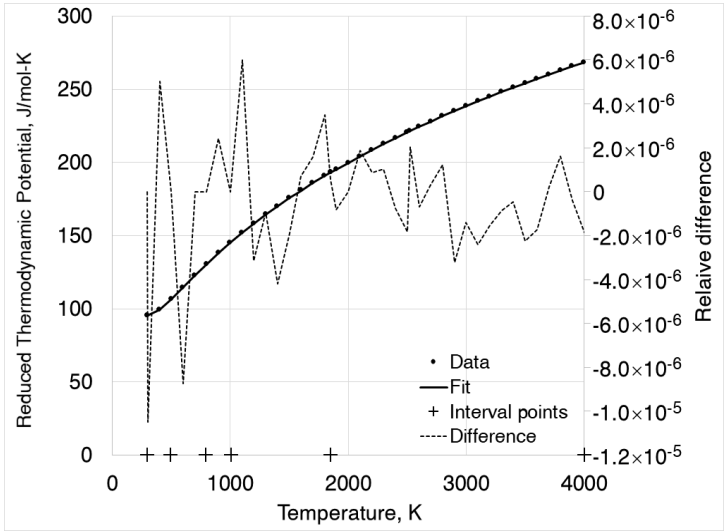
Aluminum Fluoride (AlF_3) – Al1F3(l)

MW = 83.9767 g/mol – $\Delta_f H^o$(298.15 K) = -1422.76 kJ/mol – S^o(298.15 K) = 95.379 J/mol-K – p^o = 0.1 MPa



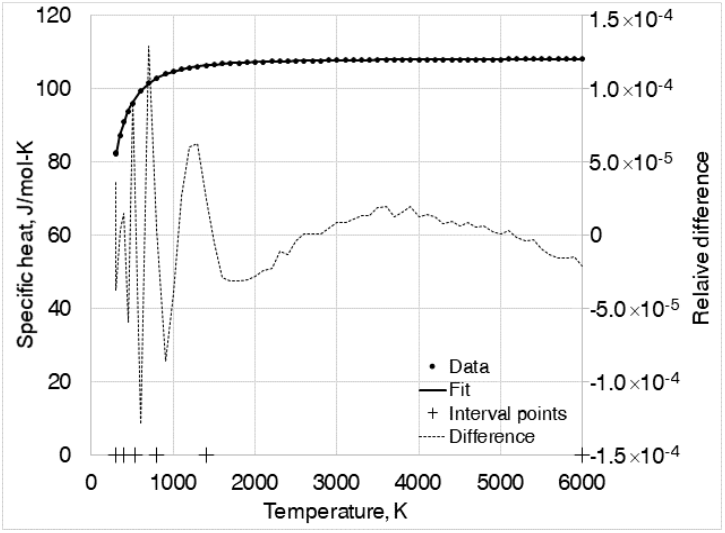


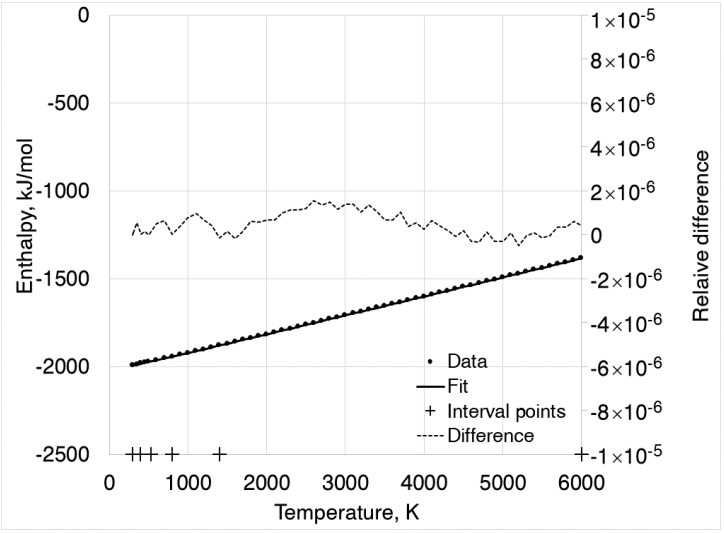


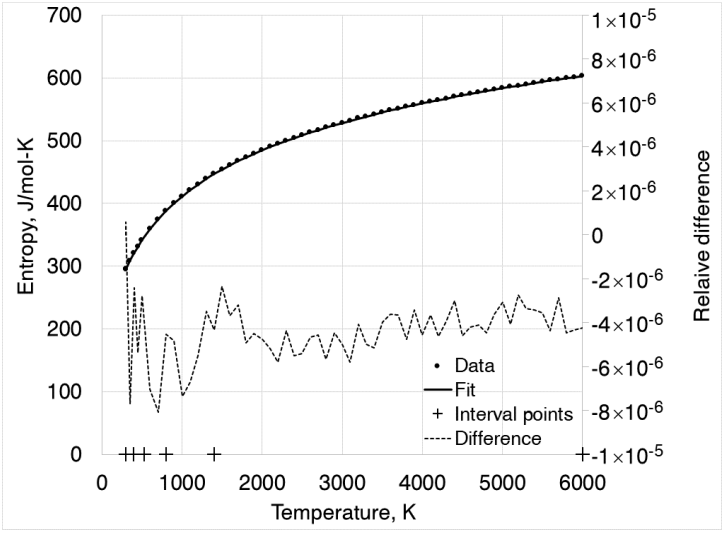


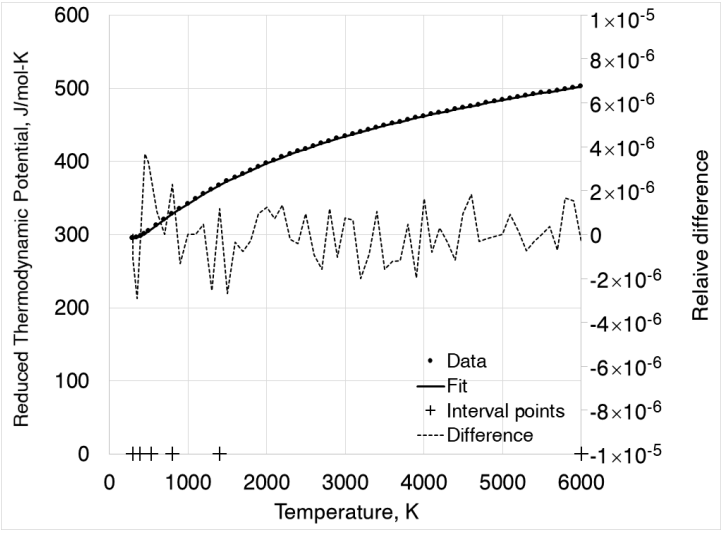
Tetrafluoroaluminate, Ion (AlF_4^-) – Al1F4-(g)

MW = 102.9757 g/mol – $\Delta_f H^o$(298.15 K) = -1991.584 kJ/mol – S^o(298.15 K) = 294.747 J/mol-K – p^o = 0.1 MPa



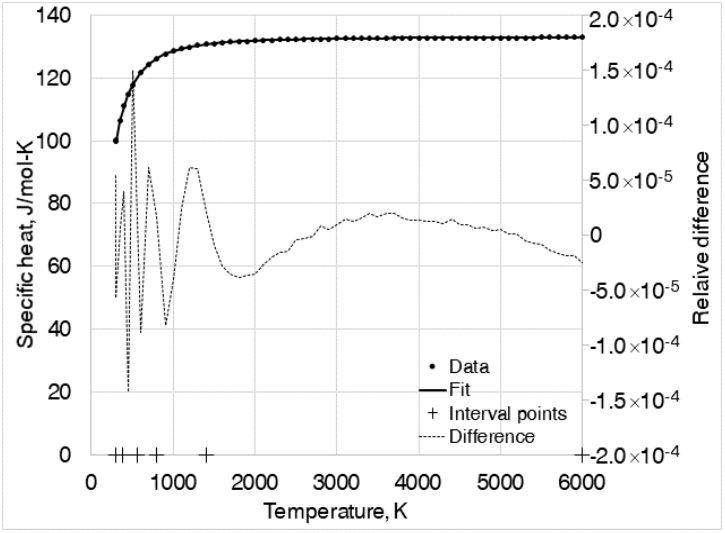


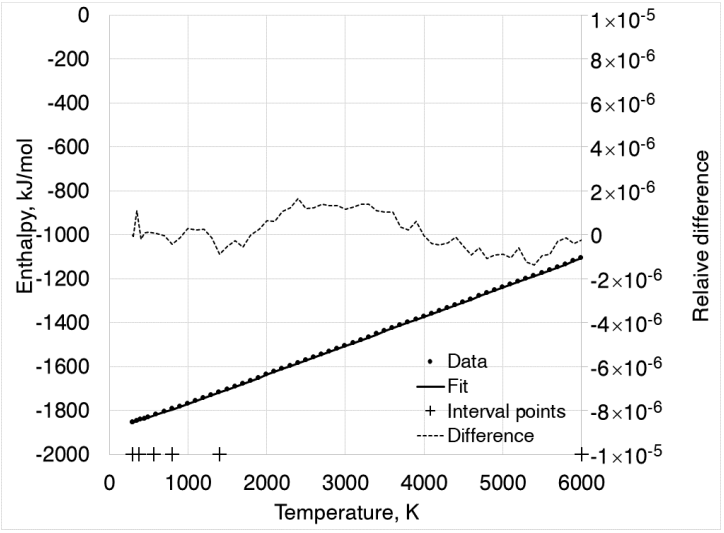


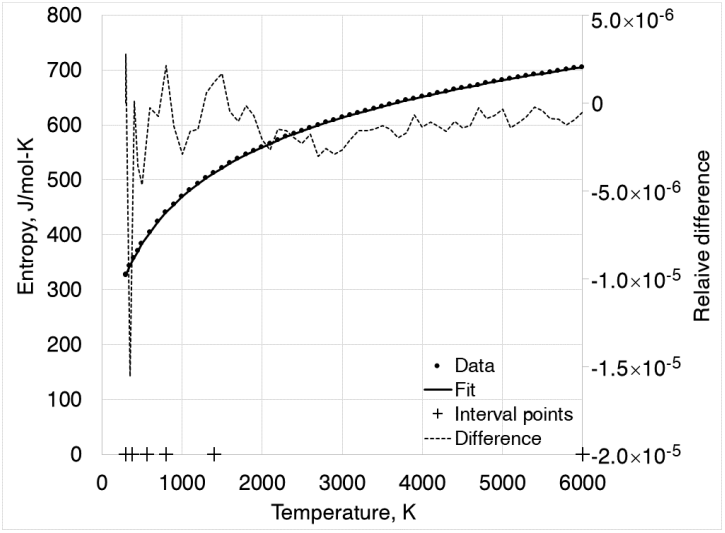


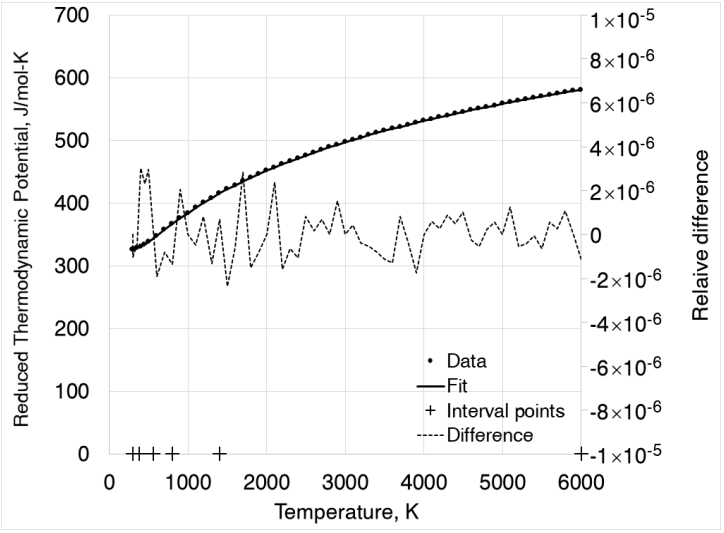
Lithium Tetrafluoroaluminate ($LiAlF_4$) – Al1F4Li1(g)

MW = 109.916 g/mol – $\Delta_f H^o$(298.15 K) = -1853.512 kJ/mol – S^o(298.15 K) = 326.497 J/mol-K – p^o = 0.1 MPa



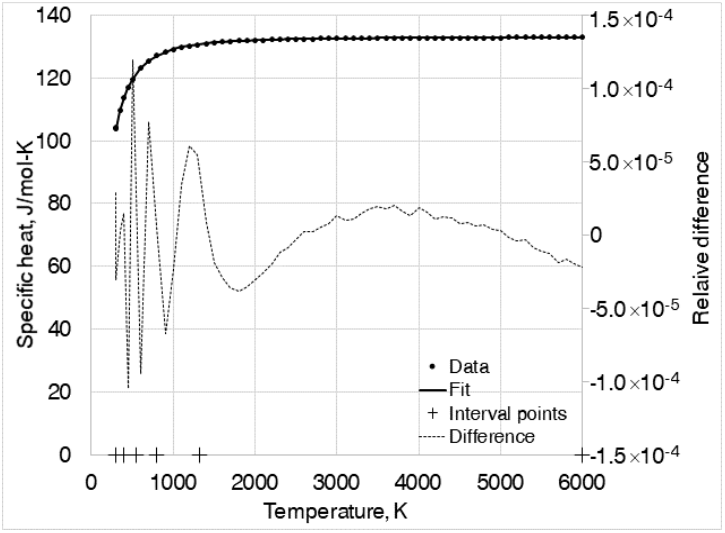


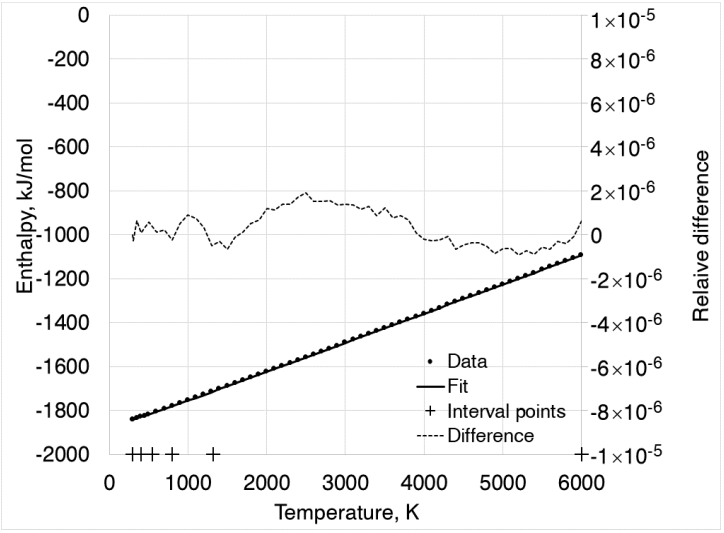


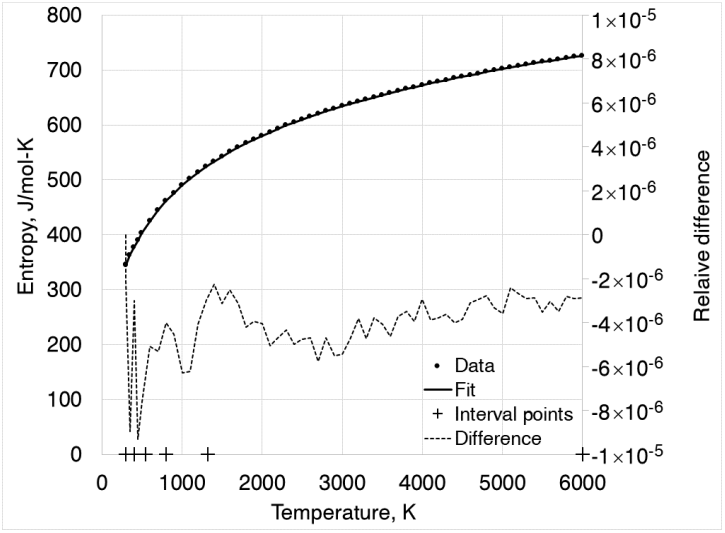


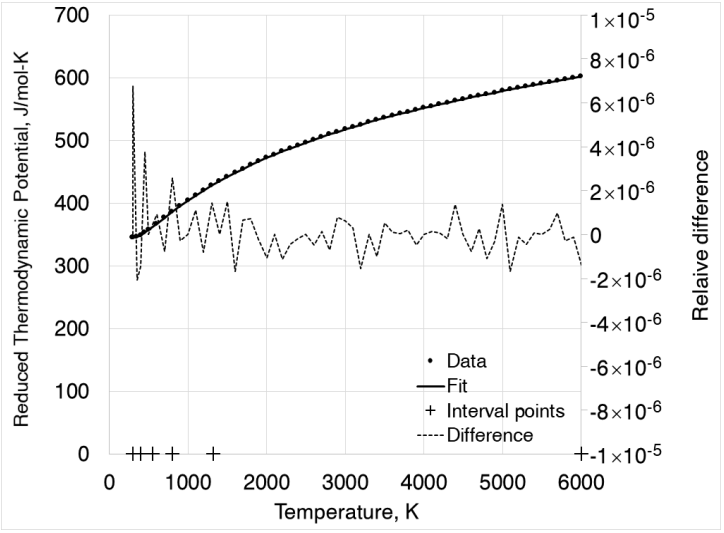
Sodium Tetrafluoroaluminate ($NaAlF_4$) – Al1F4Na1(g)

MW = 125.9649 g/mol – $\Delta_f H^o$(298.15 K) = -1840.96 kJ/mol – S^o(298.15 K) = 344.896 J/mol-K – p^o = 0.1 MPa



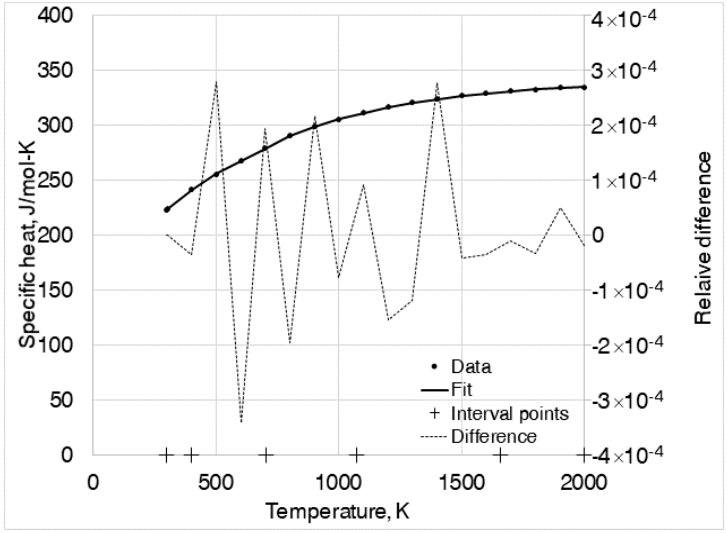


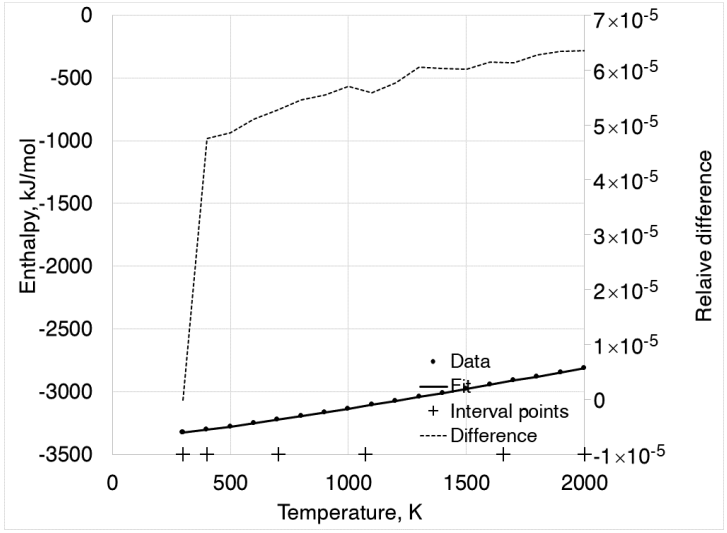


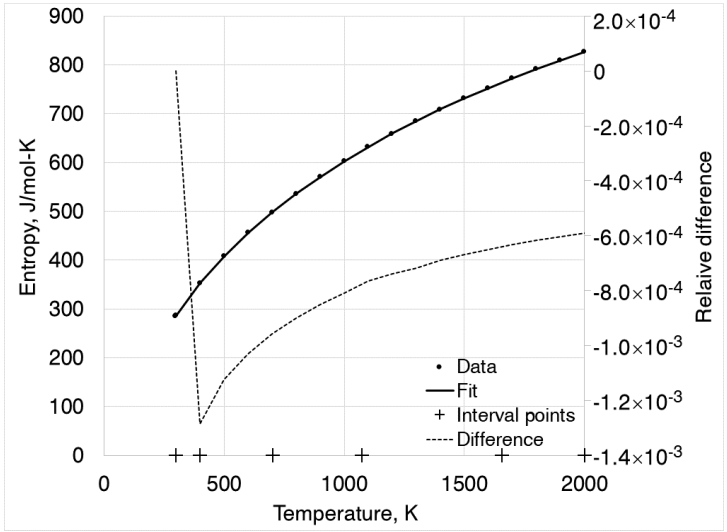


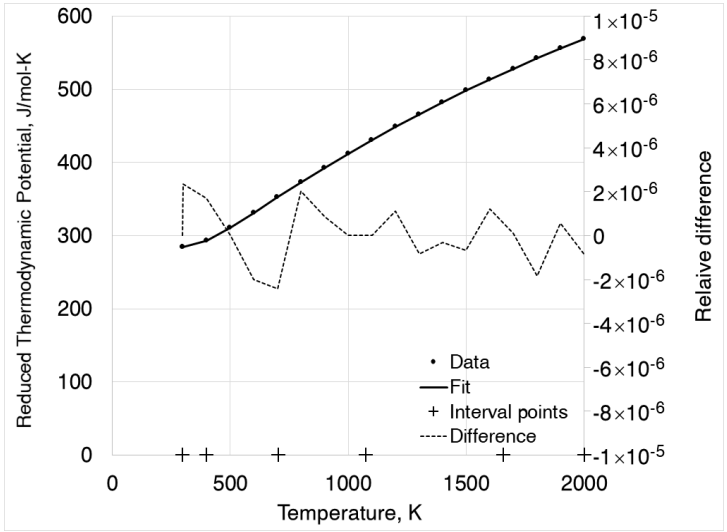
Potassium Hexafluoraluminate (K_3AlF_6) – Al1F6K3(cr)

MW = 258.2669 g/mol – $\Delta_f H^o$(298.15 K) = -3326.28 kJ/mol – S^o(298.15 K) = 284.512 J/mol-K – p^o = 0.1 MPa



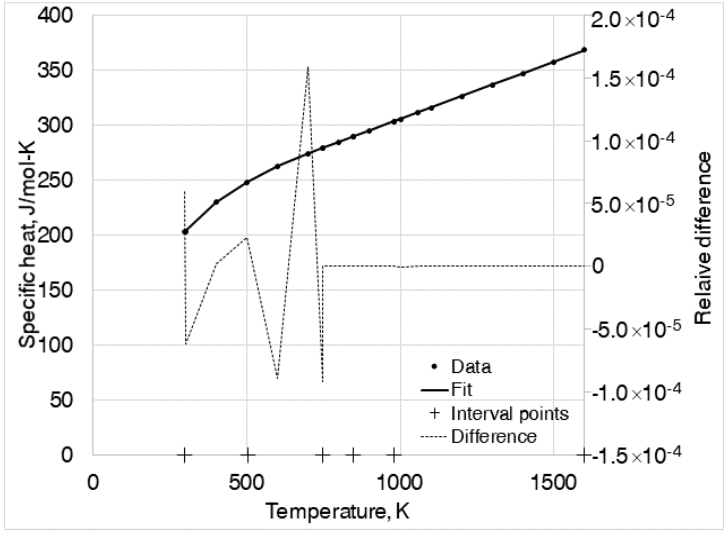


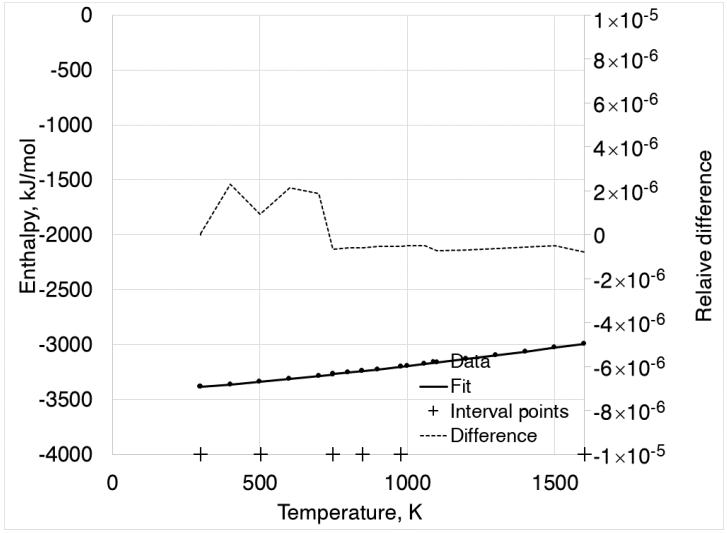


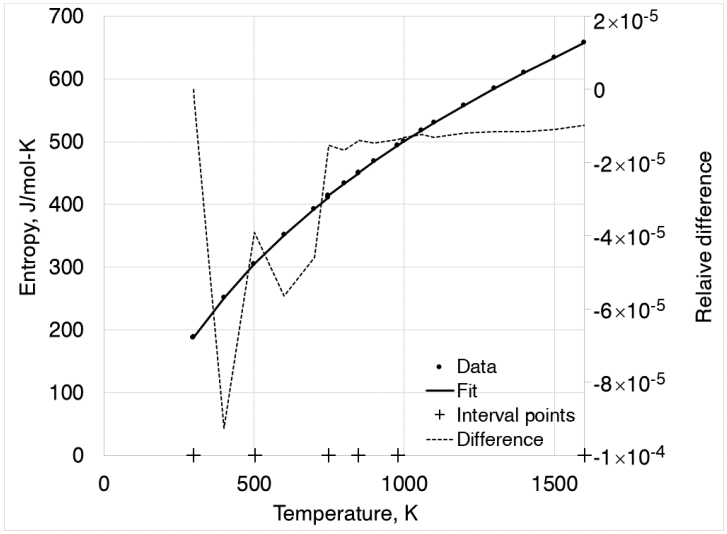


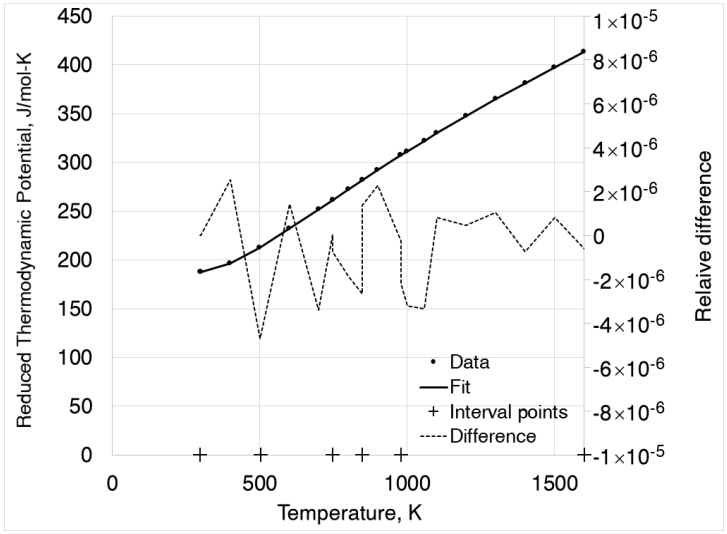
Lithium Hexafluoroaluminate (Li_3AlF_6) – Al1F6Li3(cr)

MW = 161.795 g/mol – $\Delta_f H^o$(298.15 K) = -3383.601 kJ/mol – S^o(298.15 K) = 187.891 J/mol-K – p^o = 0.1 MPa



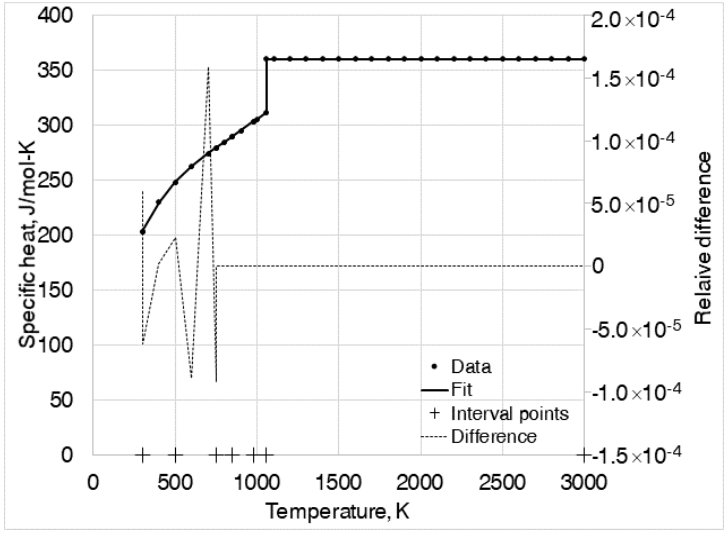


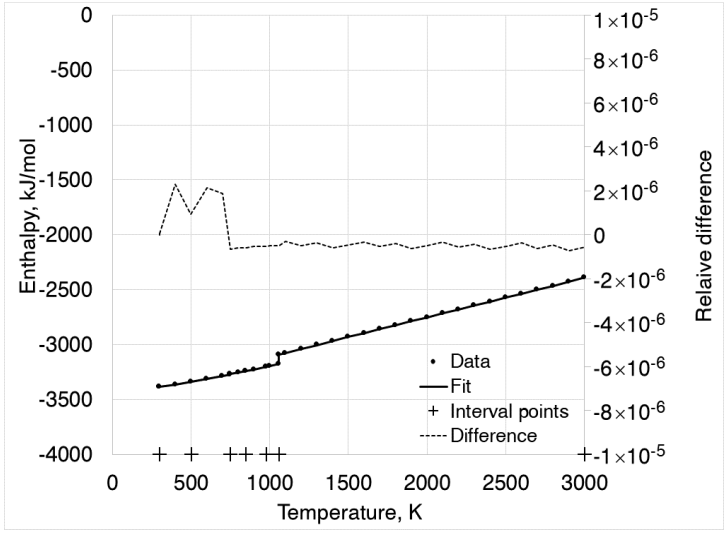


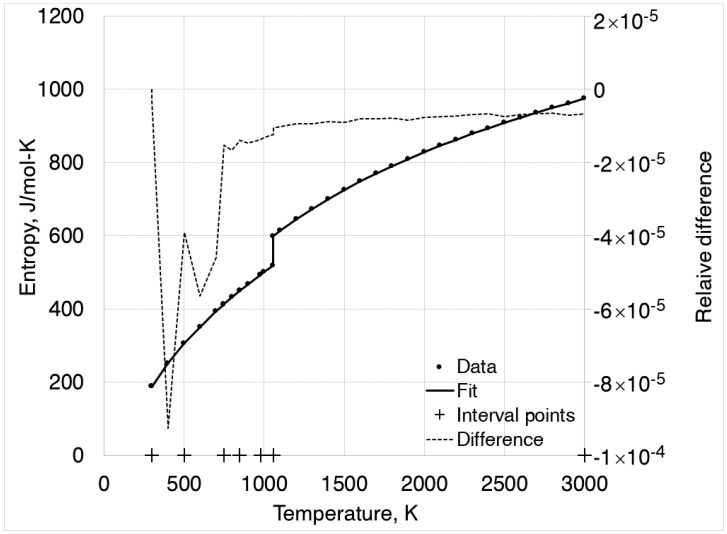


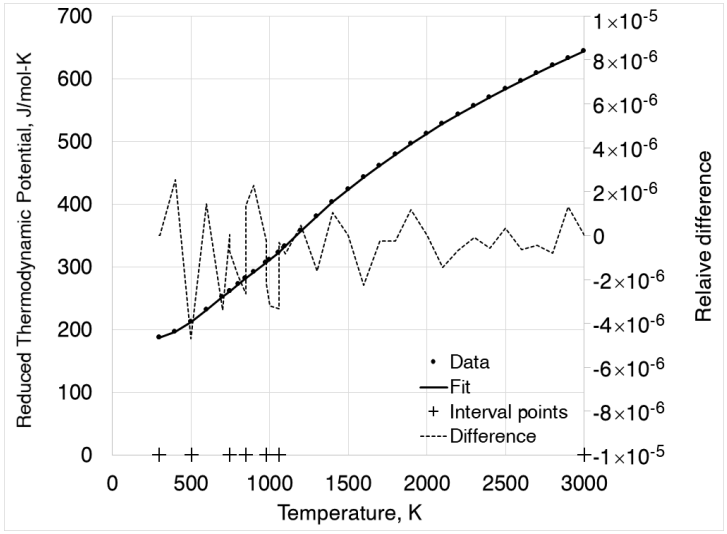
Lithium Hexafluoroaluminate (Li_3AlF_6) – Al1F6Li3(cr,l)

MW = 161.795 g/mol – $\Delta_f H^o$(298.15 K) = -3383.601 kJ/mol – S^o(298.15 K) = 187.891 J/mol-K – p^o = 0.1 MPa



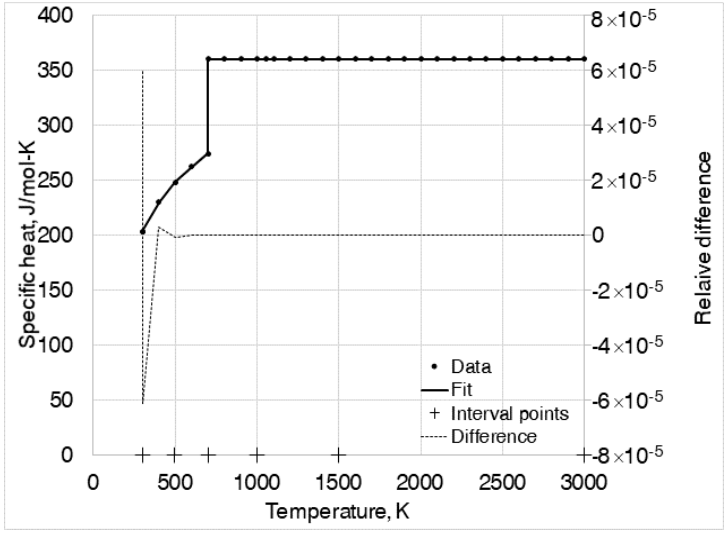


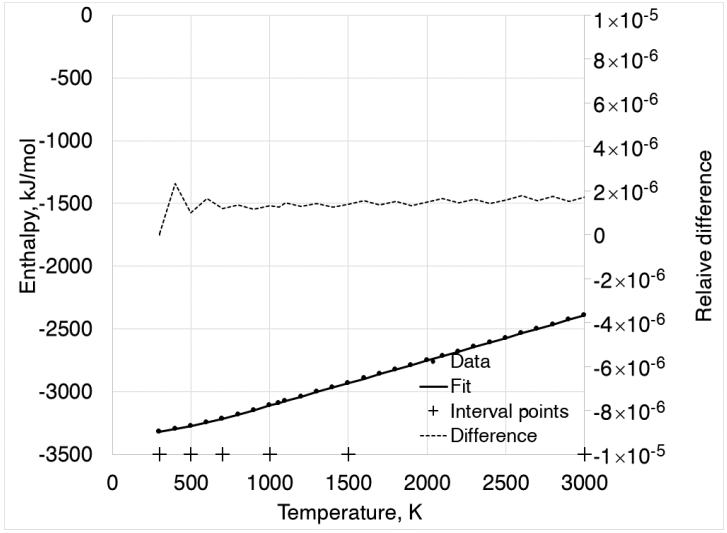


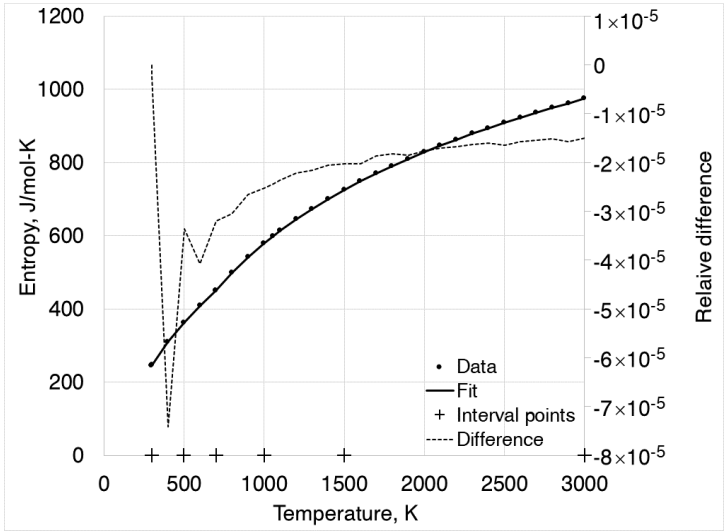


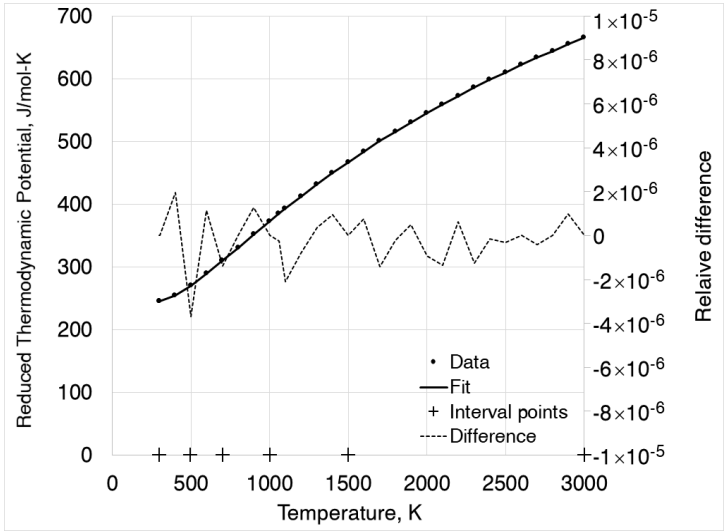
Lithium Hexafluoroaluminate (Li_3AlF_6) – Al1F6Li3(l)

MW = 161.795 g/mol – $\Delta_f H^o$(298.15 K) = -3317.64 kJ/mol – S^o(298.15 K) = 245.847 J/mol-K – p^o = 0.1 MPa



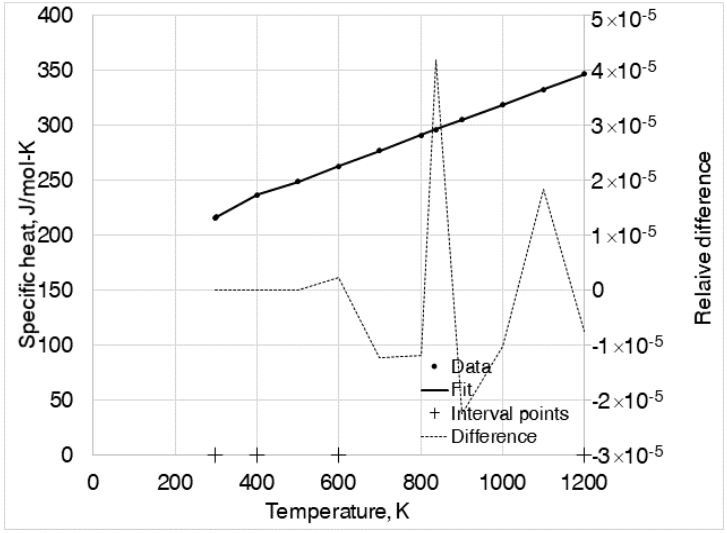


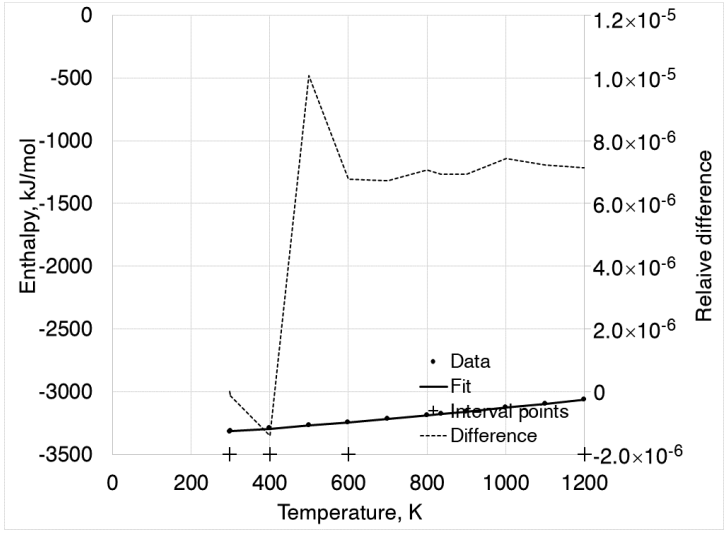


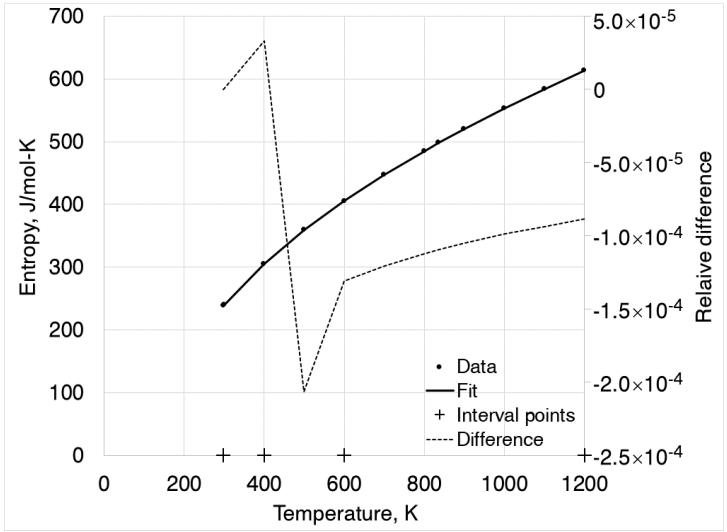


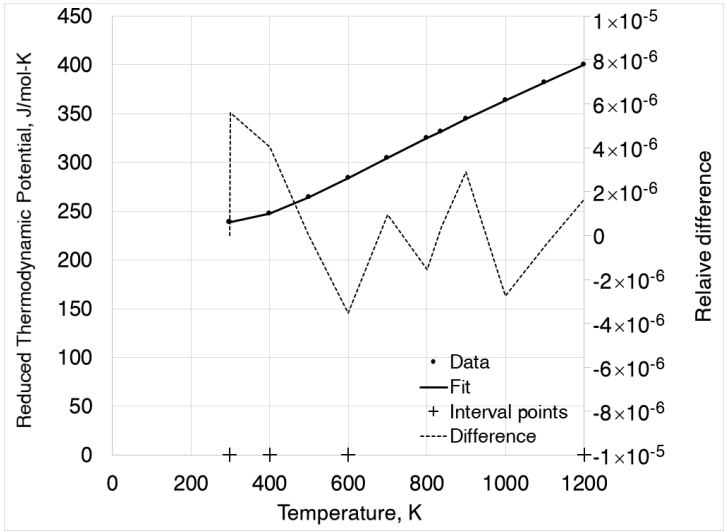
Cryolite, Alpha (Na_3AlF_6) – Al1F6Na3(cr)_01

MW = 209.9413 g/mol – $\Delta_f H^o$(298.15 K) = -3316.908 kJ/mol – S^o(298.15 K) = 238.467 J/mol-K – p^o = 0.1 MPa



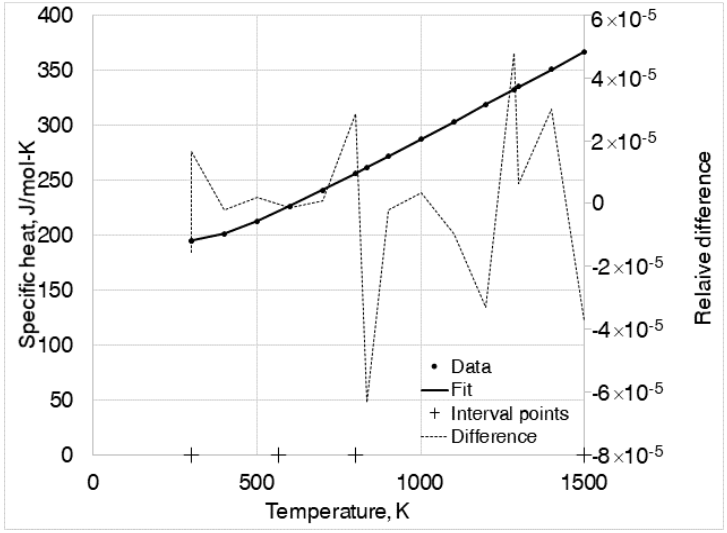


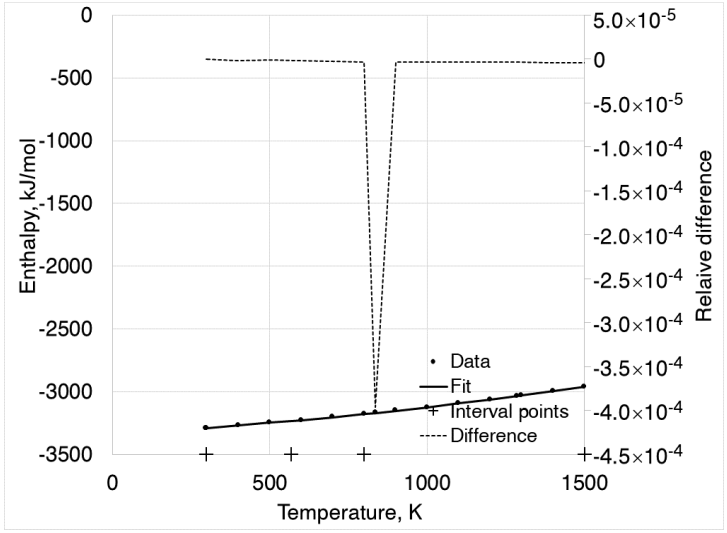


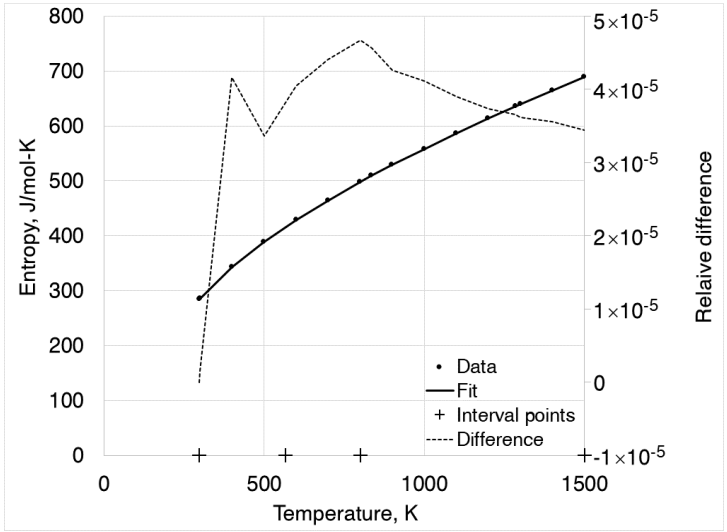


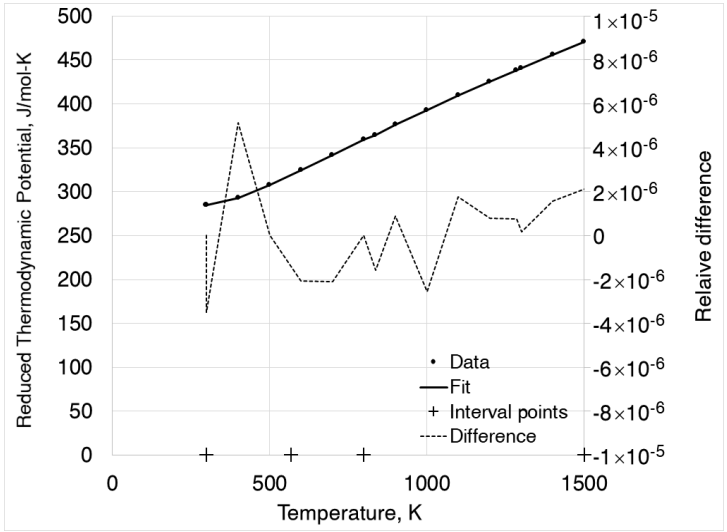
Cryolite, Beta (Na_3AlF_6) – Al1F6Na3(cr)_02

MW = 209.9413 g/mol – $\Delta_f H^o$(298.15 K) = -3288.781 kJ/mol – S^o(298.15 K) = 284.881 J/mol-K – p^o = 0.1 MPa



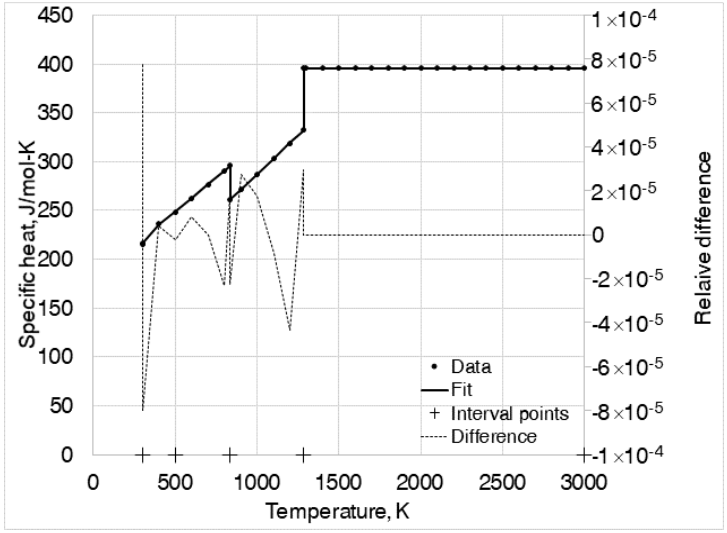


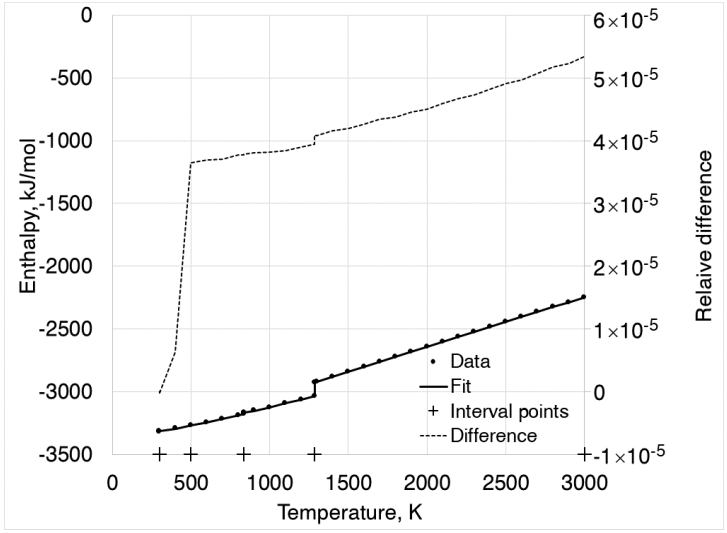


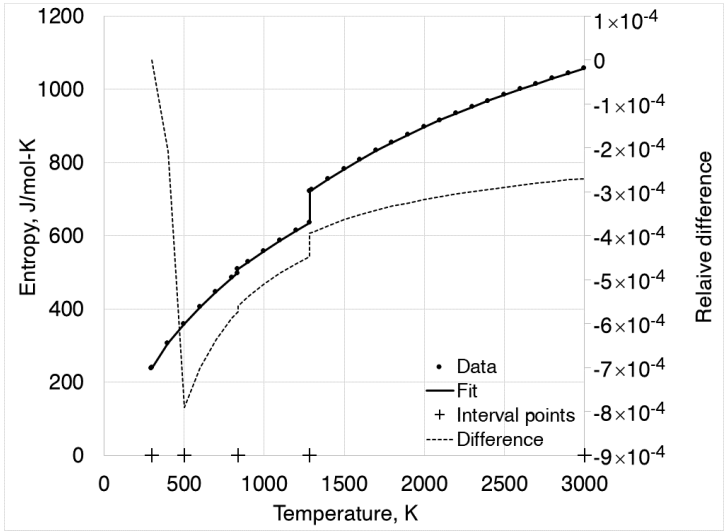


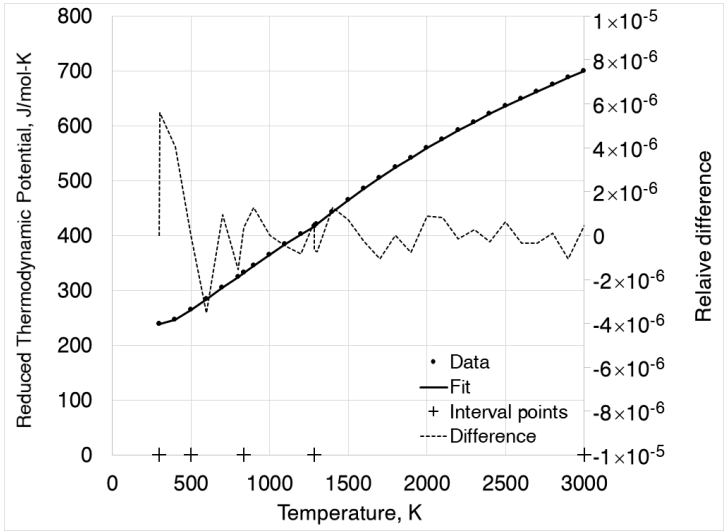
Cryolite (Na_3AlF_6) – Al1F6Na3(cr,l)

MW = 209.9413 g/mol – $\Delta_f H^o$(298.15 K) = -3316.908 kJ/mol – S^o(298.15 K) = 238.467 J/mol-K – p^o = 0.1 MPa



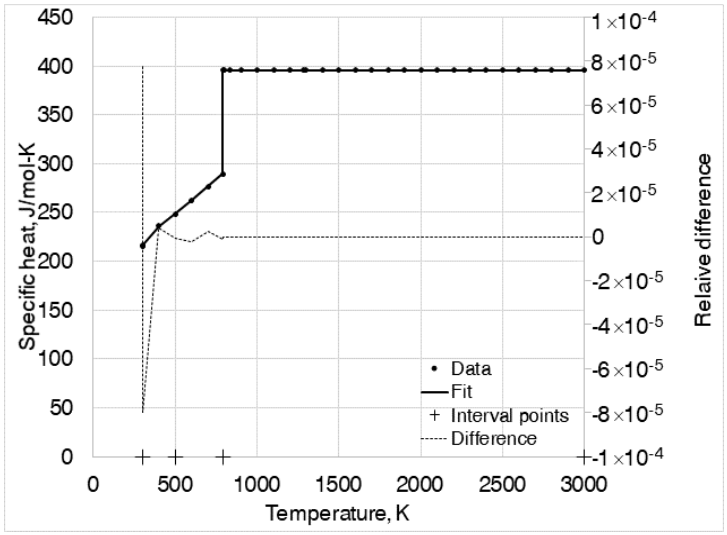


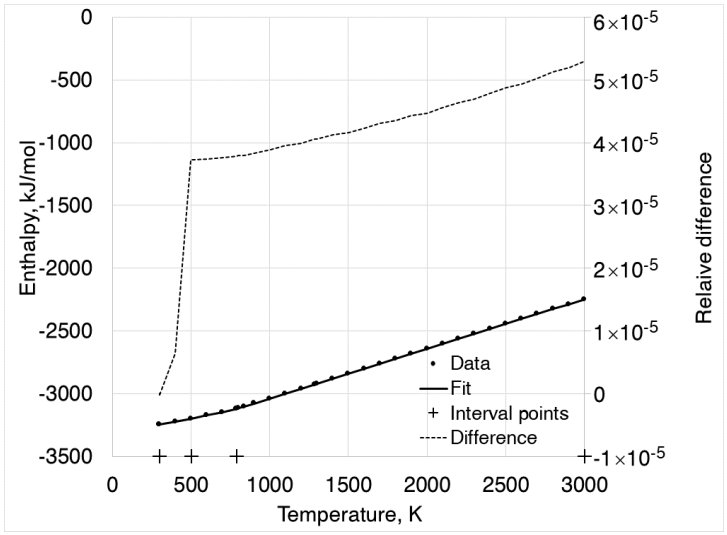


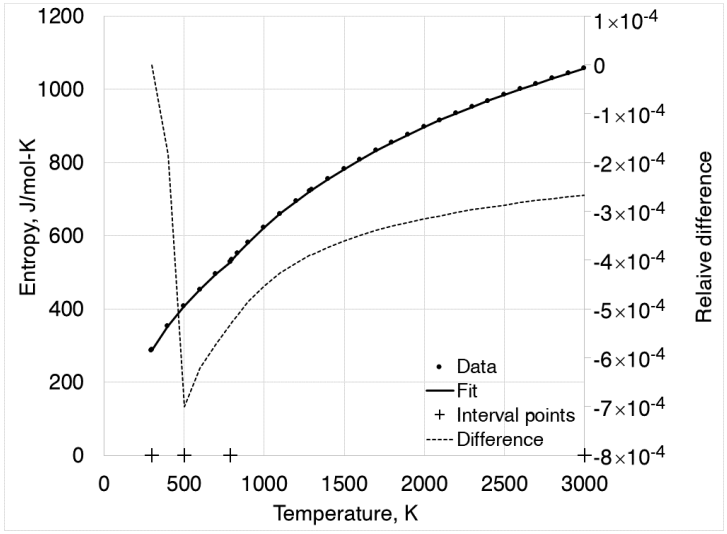


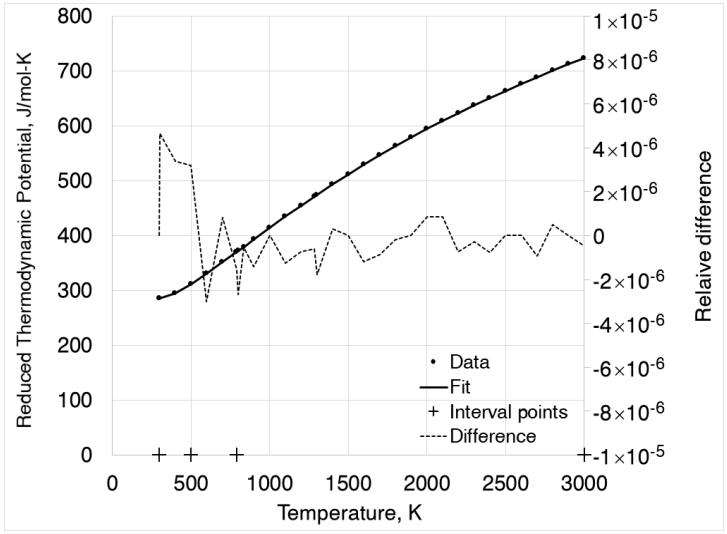
Cryolite (Na_3AlF_6) – Al1F6Na3(l)

MW = 209.9413 g/mol – $\Delta_f H^o$(298.15 K) = -3246.472 kJ/mol – S^o(298.15 K) = 286.058 J/mol-K – p^o = 0.1 MPa



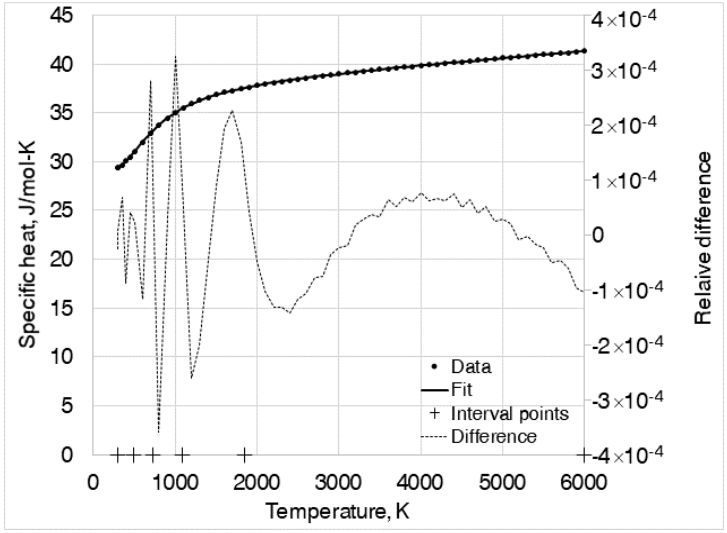


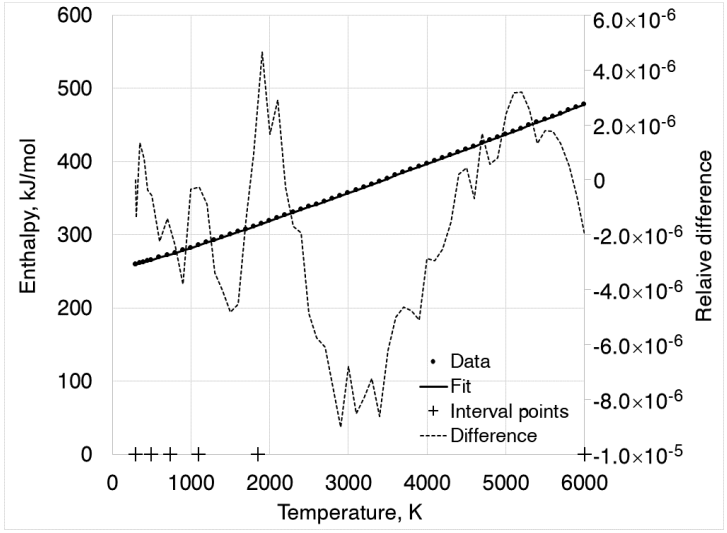


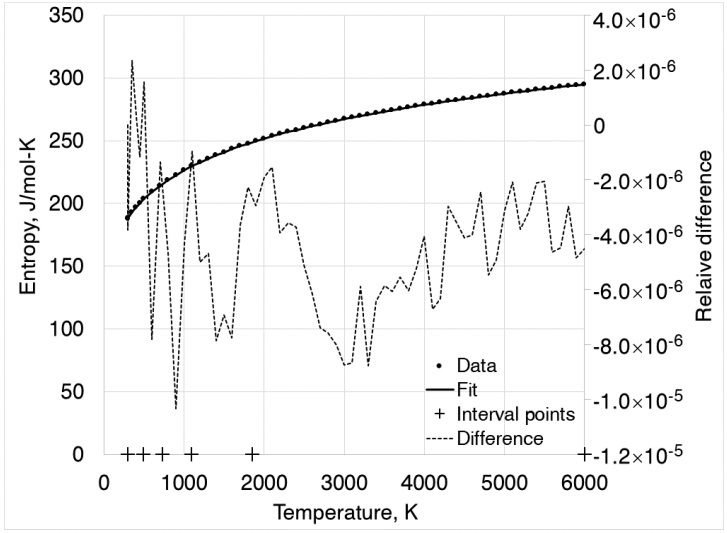


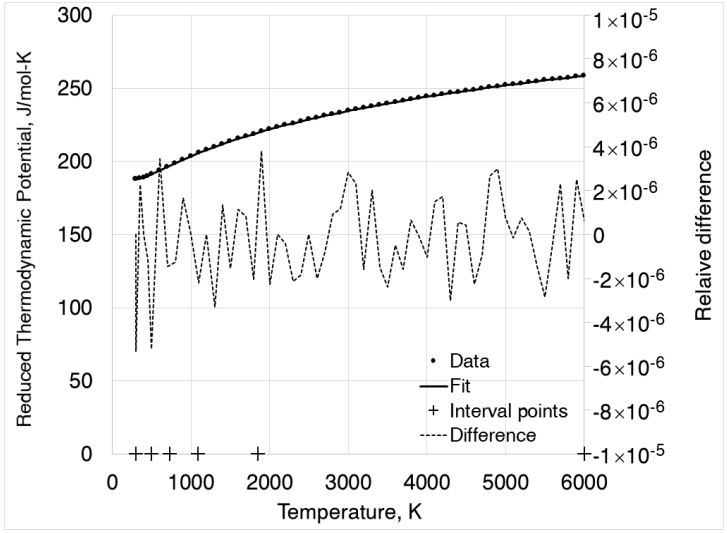
Aluminum Hydride (AlH) – Al1H1(g)

MW = 27.98948 g/mol – $\Delta_f H^o$(298.15 K) = 259.408 kJ/mol – S^o(298.15 K) = 187.864 J/mol-K – p^o = 0.1 MPa



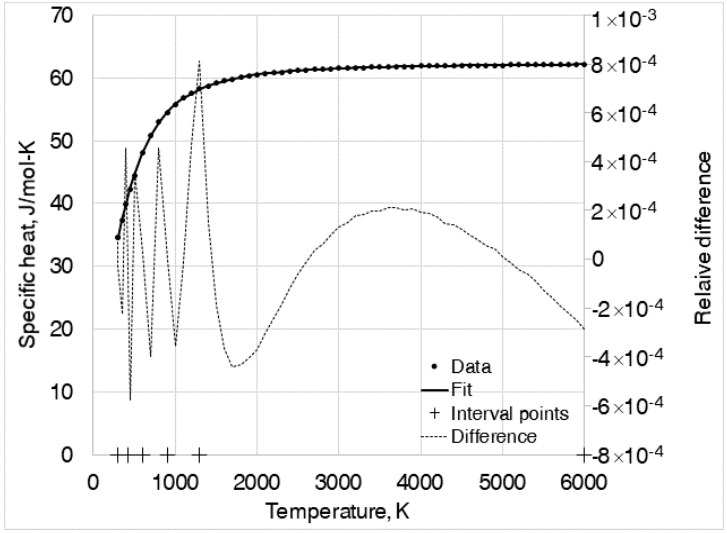


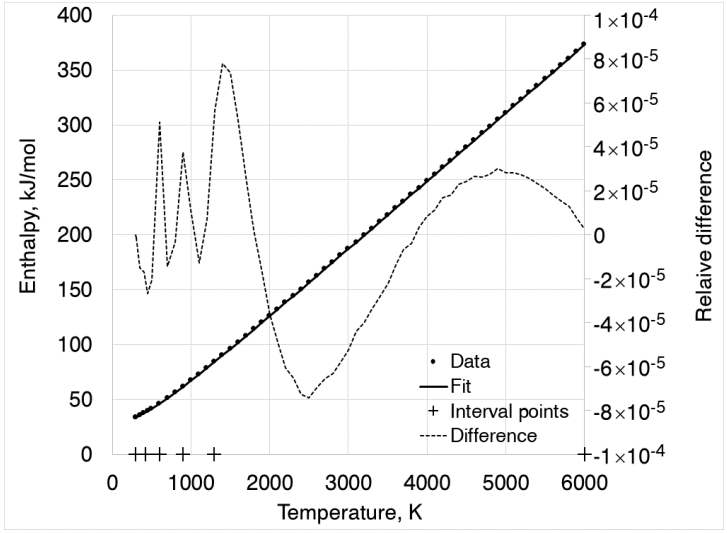


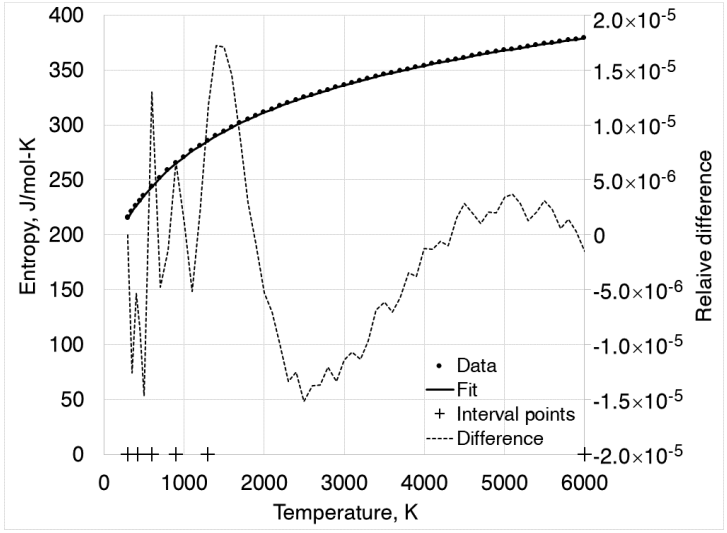


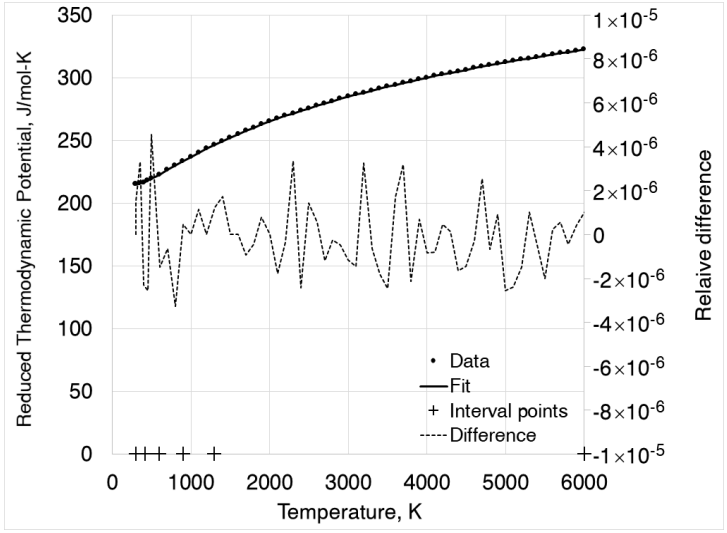
Aluminum Hydride Oxide (OAlH) – Al1H1O1(g)_01

MW = 43.9889 g/mol – $\Delta_f H^o$(298.15 K) = 33.472 kJ/mol – S^o(298.15 K) = 215.256 J/mol-K – p^o = 0.1 MPa



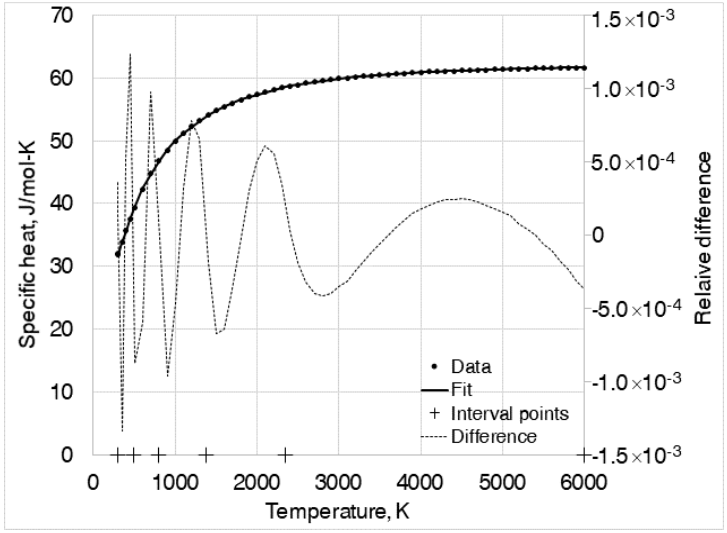


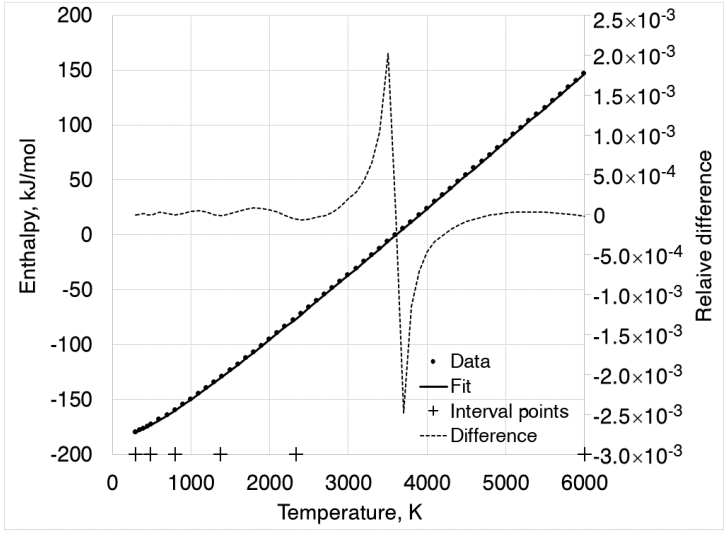


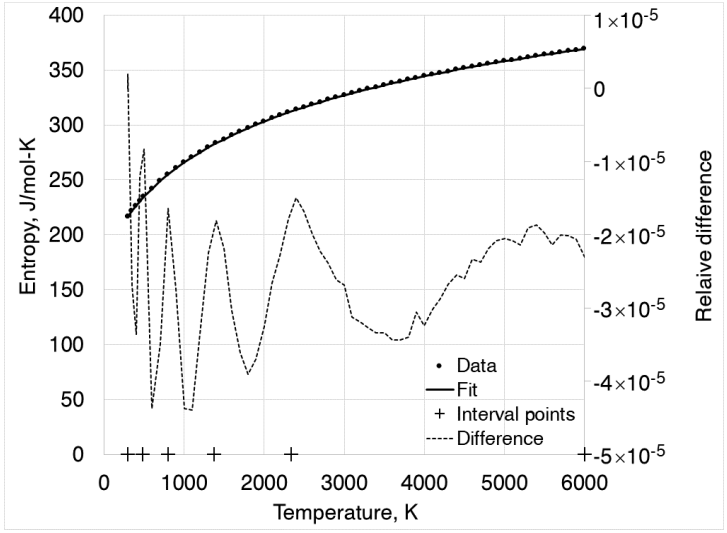


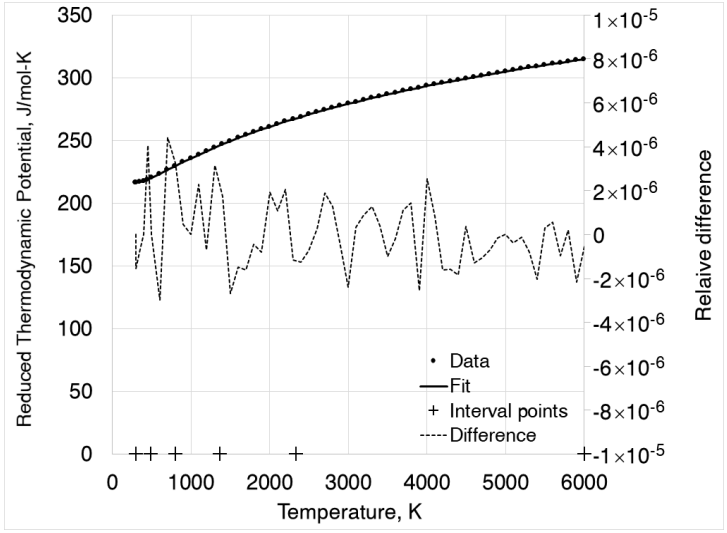
Aluminum Hydroxide (AlOH) – Al1H1O1(g)_02

MW = 43.9889 g/mol – $\Delta_f H^o$(298.15 K) = -179.912 kJ/mol – S^o(298.15 K) = 216.43 J/mol-K – p^o = 0.1 MPa



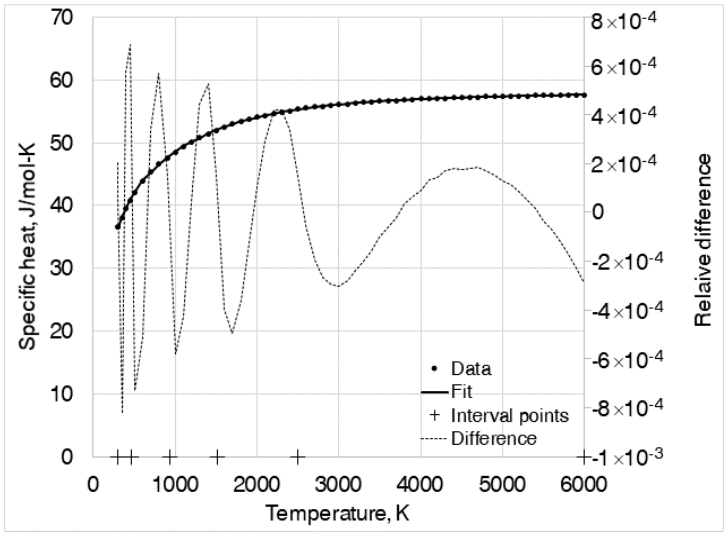


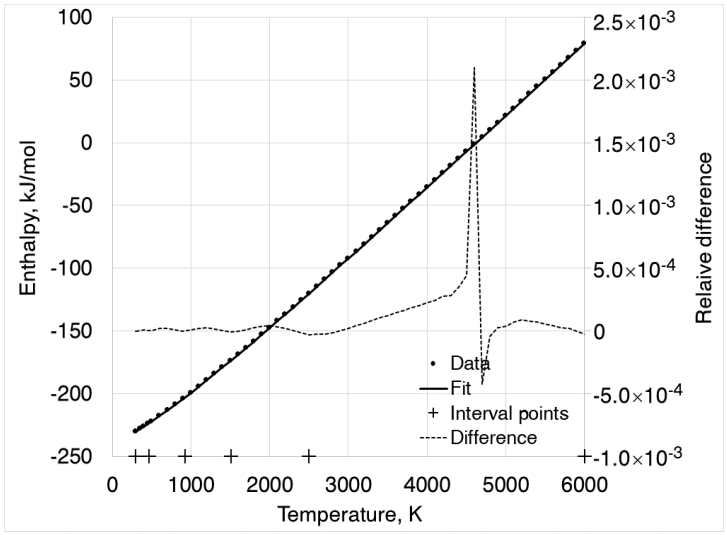


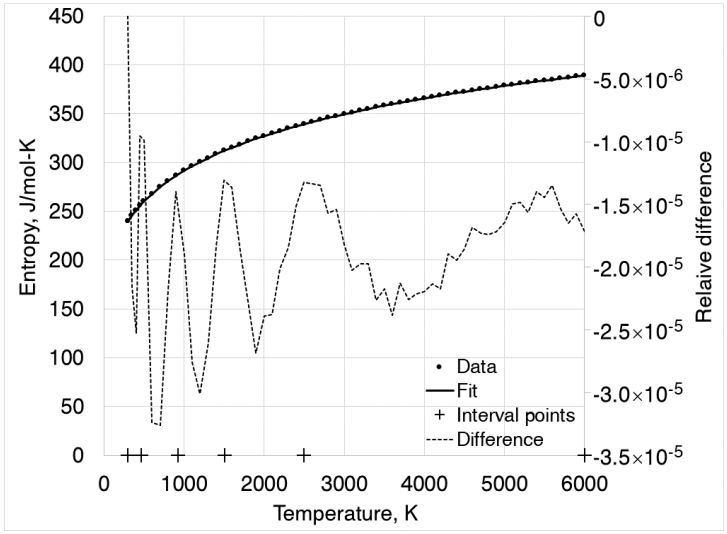


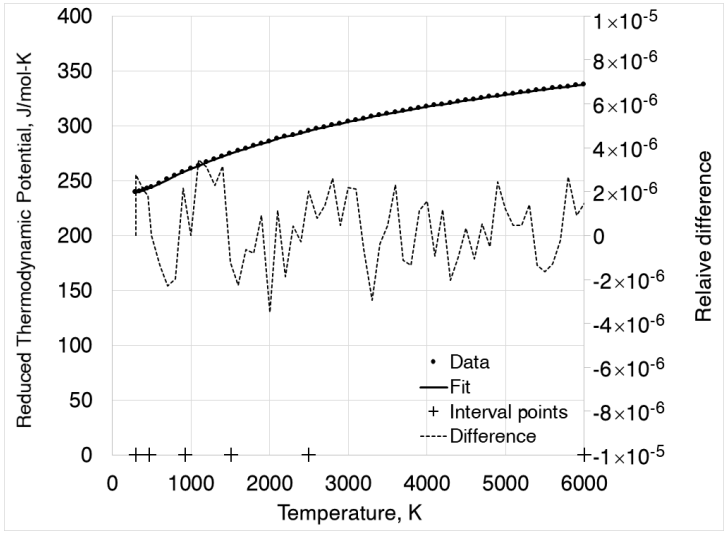
Aluminum Hydroxide, Ion ($AlOH^-$) – Al1H1O1-(g)

MW = 43.9894 g/mol – $\Delta_f H^o$(298.15 K) = -230.12 kJ/mol – S^o(298.15 K) = 239.81 J/mol-K – p^o = 0.1 MPa



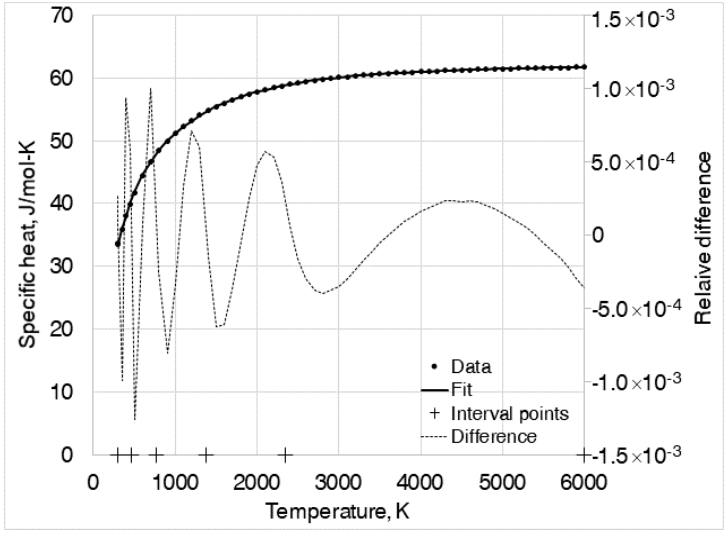


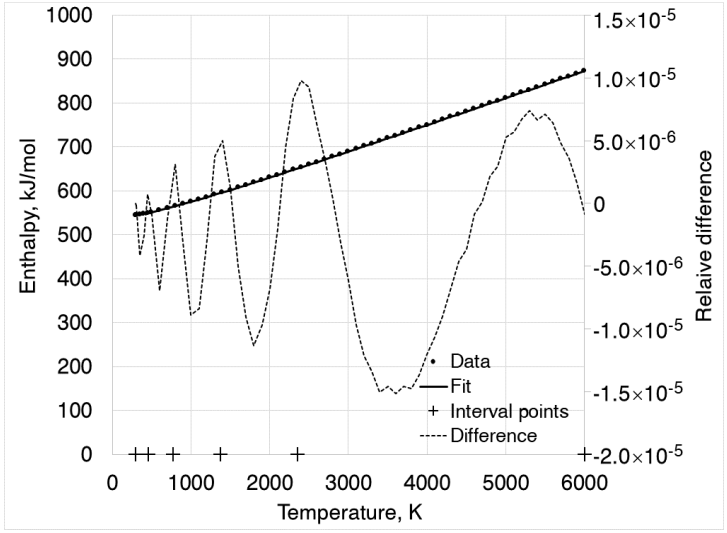


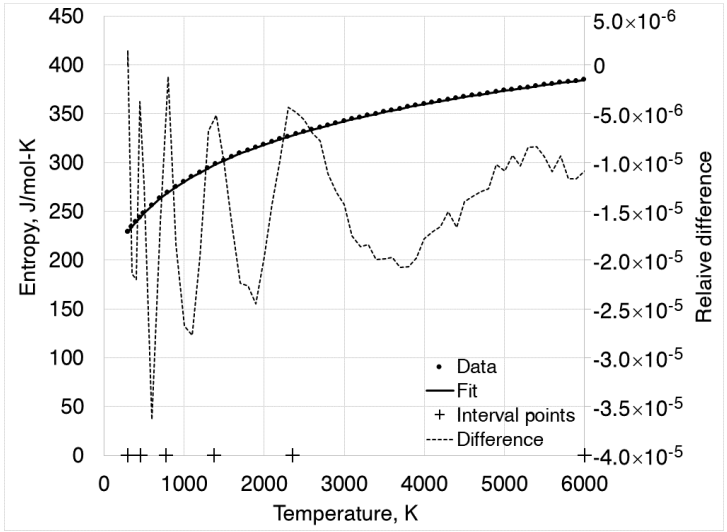


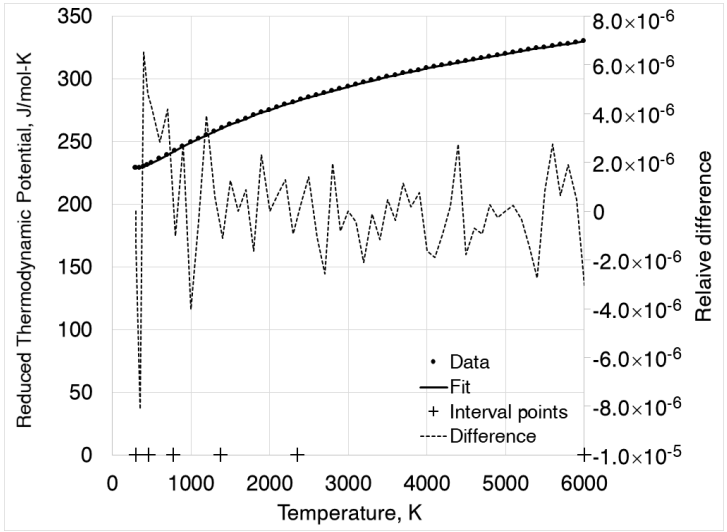
Aluminum Hydroxide, Ion ($AlOH^+$) – Al1H1O1+(g)

MW = 43.9883 g/mol – $\Delta_f H^o$(298.15 K) = 543.92 kJ/mol – S^o(298.15 K) = 228.693 J/mol-K – p^o = 0.1 MPa



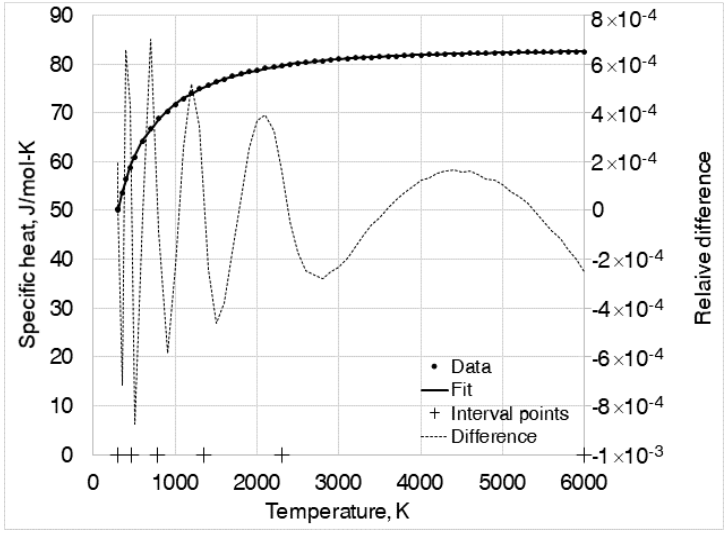


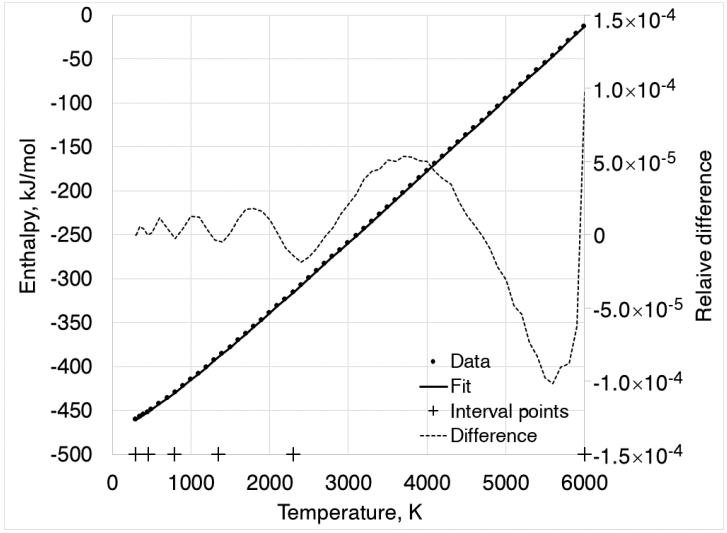


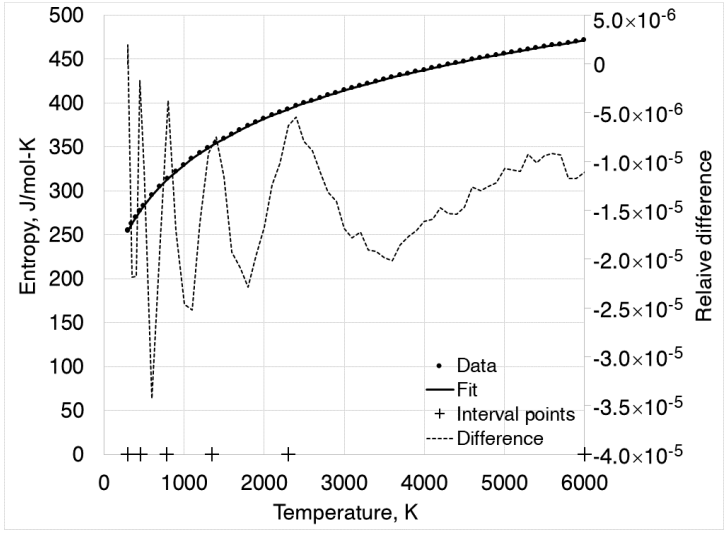


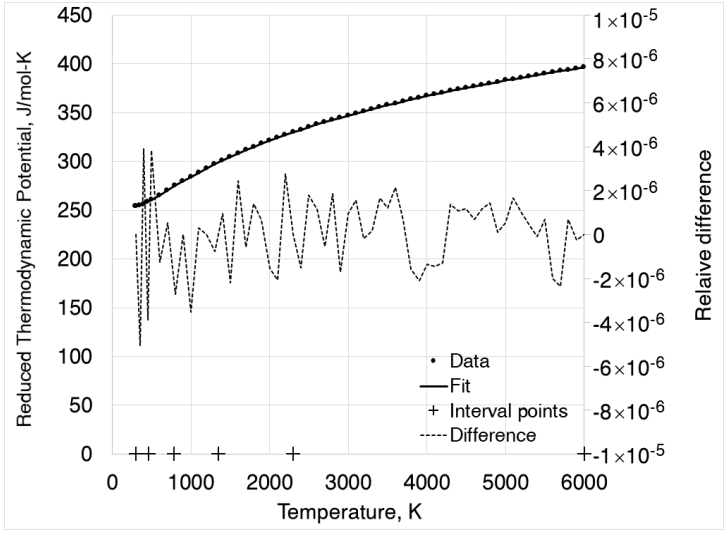
Aluminum Hydroxide Oxide (OAlOH) – Al1H1O2(g)

MW = 59.9883 g/mol – $\Delta_f H^o$(298.15 K) = -460.24 kJ/mol – S^o(298.15 K) = 254.391 J/mol-K – p^o = 0.1 MPa



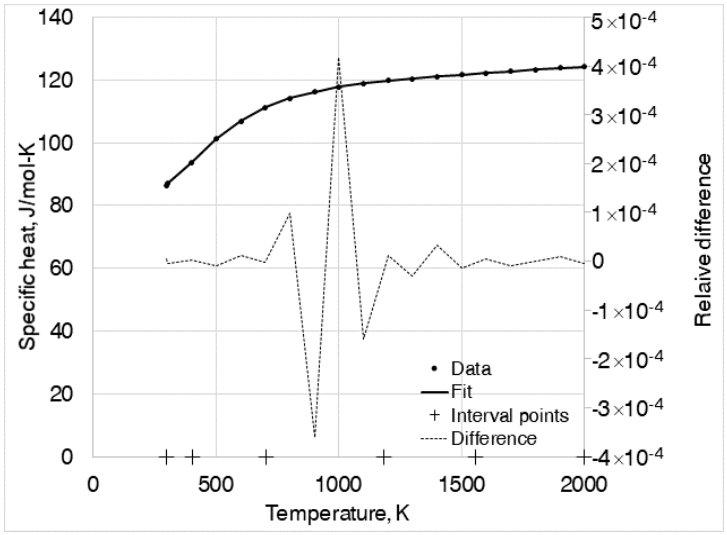


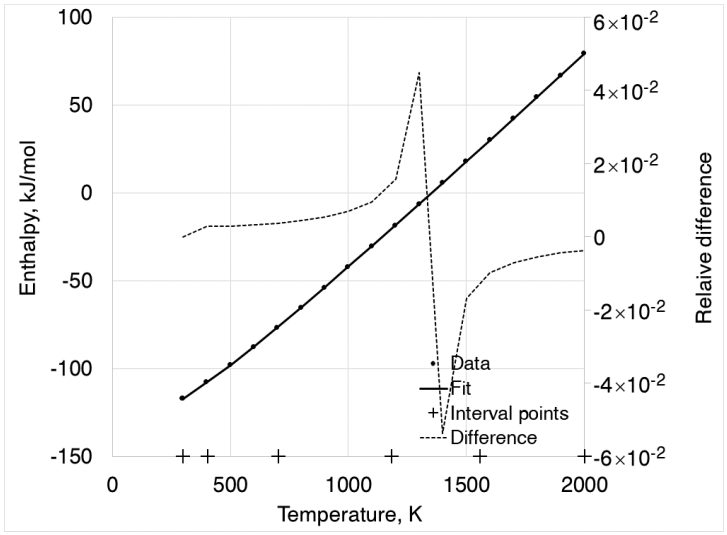


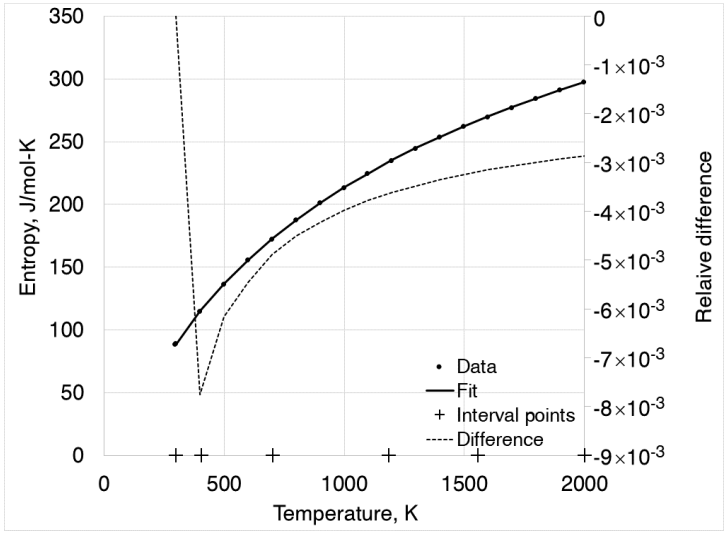


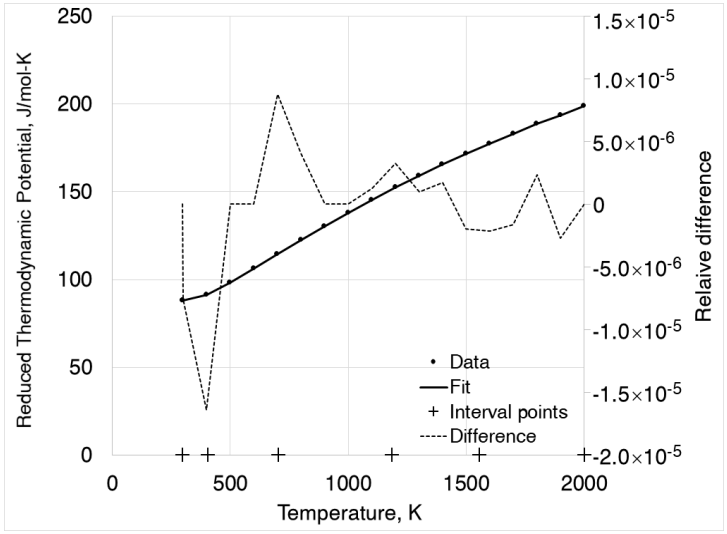
Lithium Tetrahydroaluminate ($LiAlH_4$) – Al1H4Li1(cr)

MW = 37.954 g/mol – $\Delta_f H^o$(298.15 K) = -117.152 kJ/mol – S^o(298.15 K) = 87.864 J/mol-K – p^o = 0.1 MPa



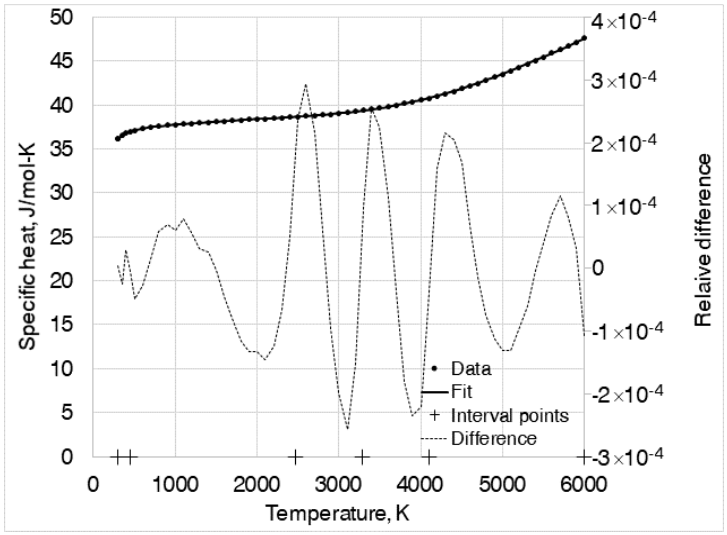


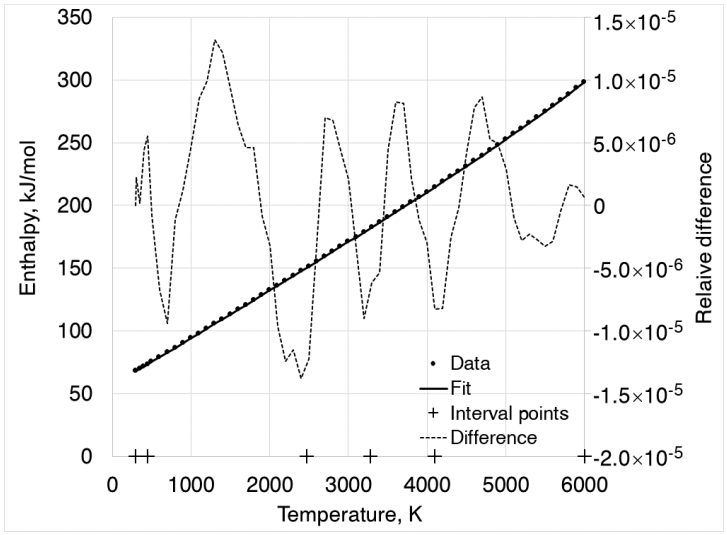


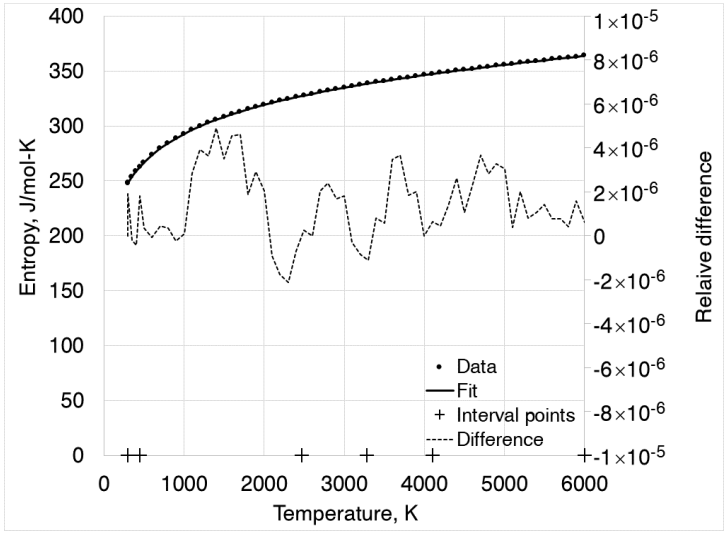


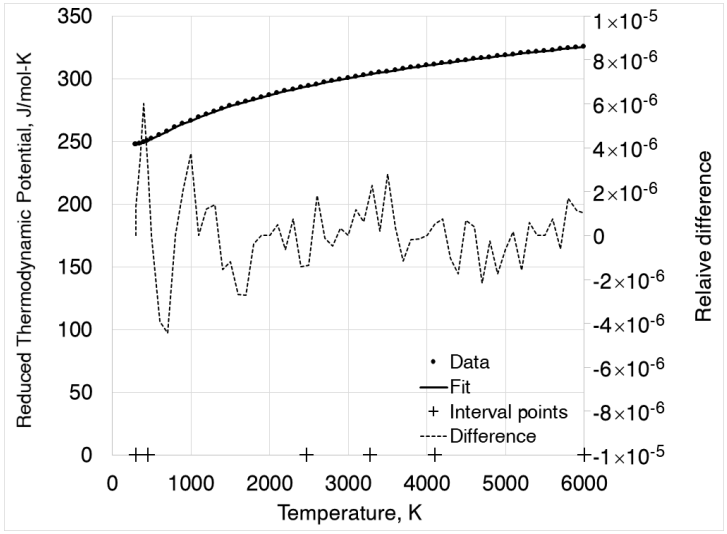
Aluminum Iodide (AlI) – Al1I1(g)

MW = 153.886 g/mol – $\Delta_f H^o$(298.15 K) = 67.948 kJ/mol – S^o(298.15 K) = 247.83 J/mol-K – p^o = 0.1 MPa



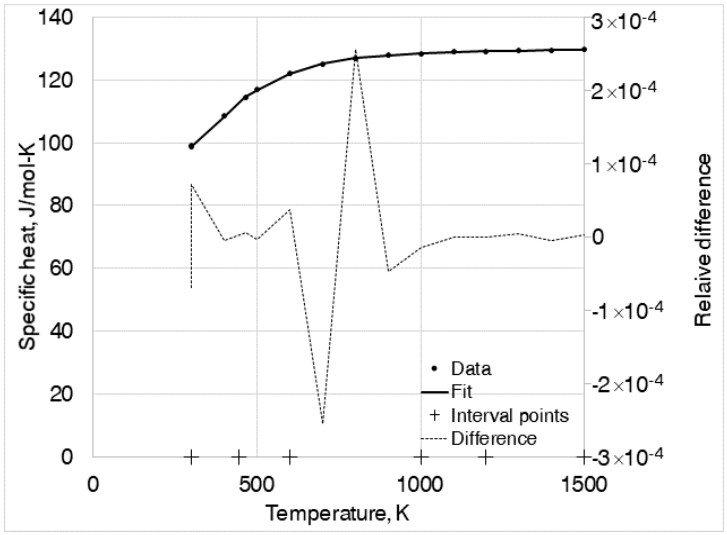


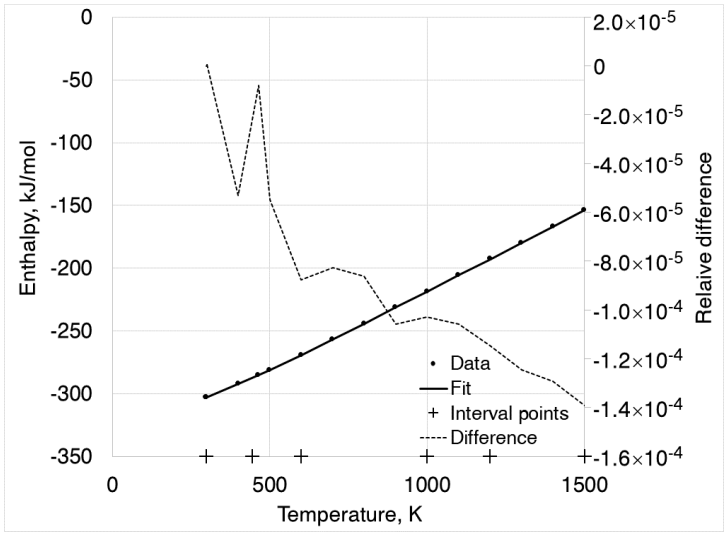


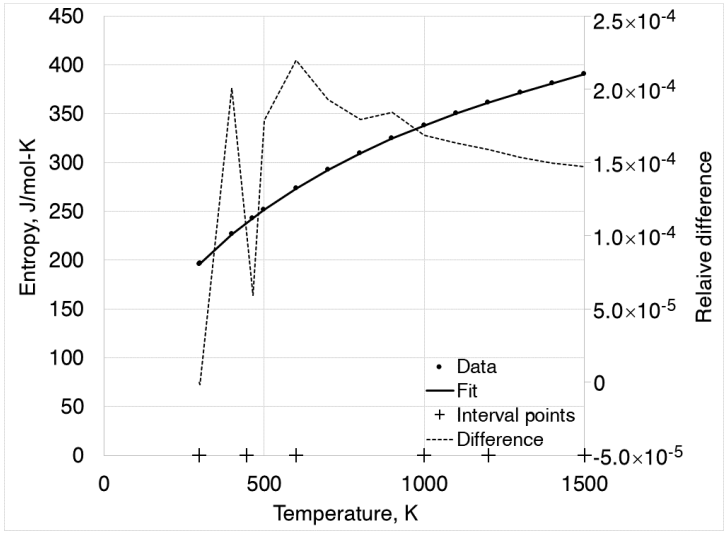


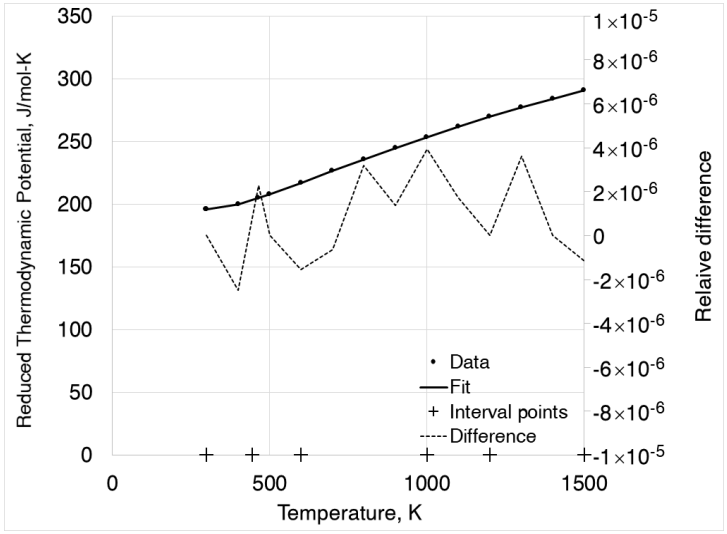
Aluminum Iodide (AlI_3) – Al1I3(cr)

MW = 407.695 g/mol – $\Delta_f H^o$(298.15 K) = -302.922 kJ/mol – S^o(298.15 K) = 196.104 J/mol-K – p^o = 0.1 MPa



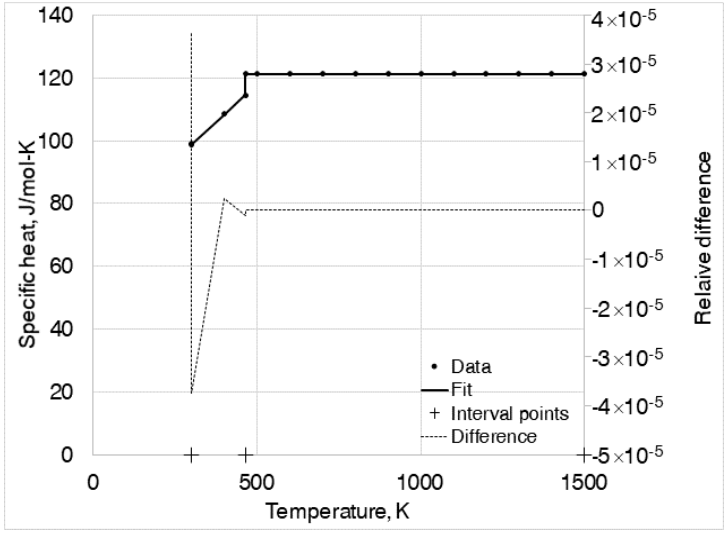


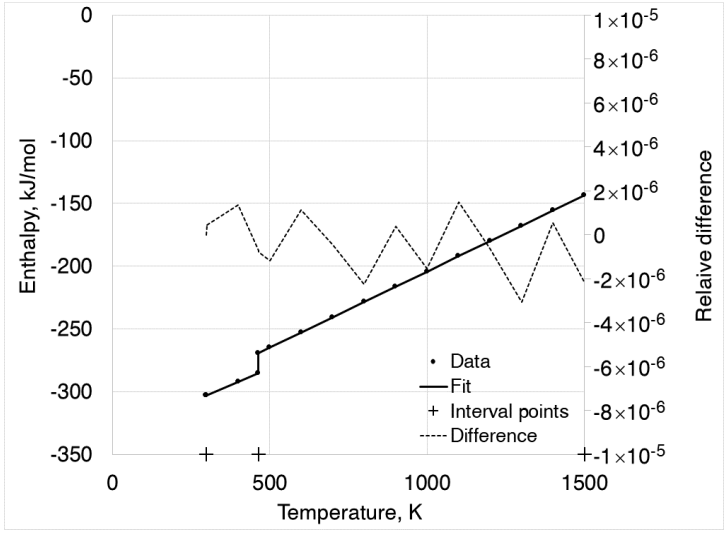


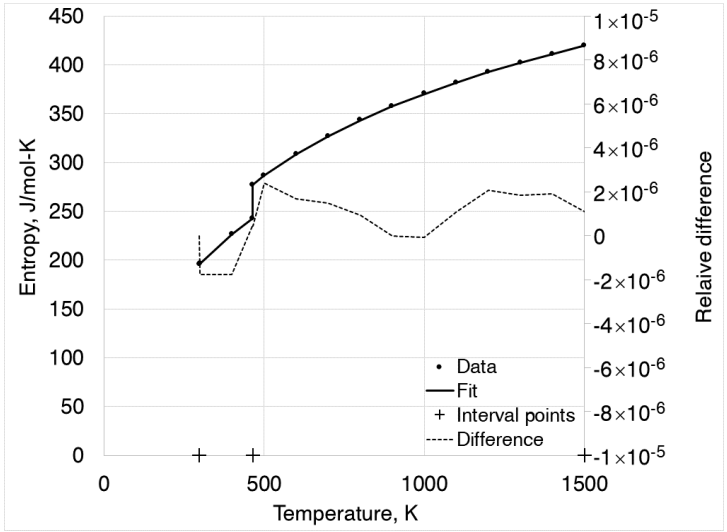


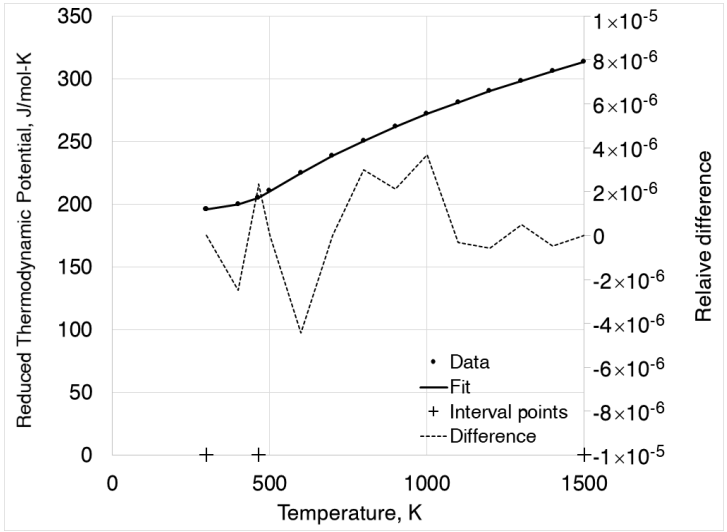
Aluminum Iodide (AlI_3) – Al1I3(cr,l)

MW = 407.695 g/mol – $\Delta_f H^o$(298.15 K) = -302.922 kJ/mol – S^o(298.15 K) = 196.104 J/mol-K – p^o = 0.1 MPa



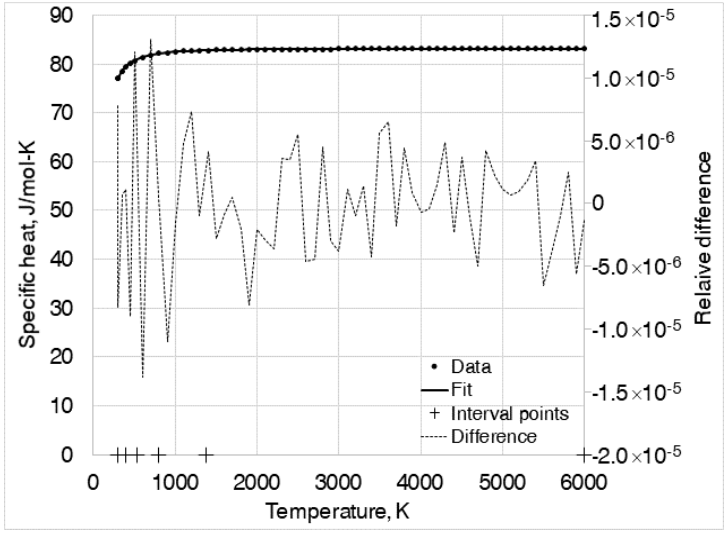


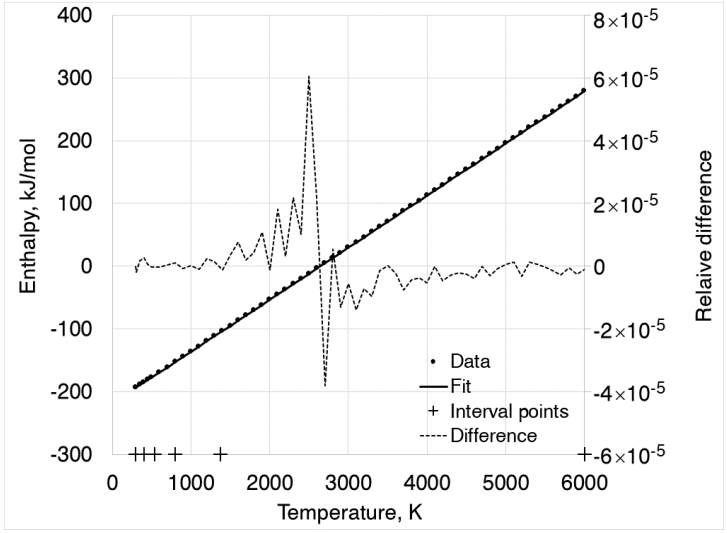


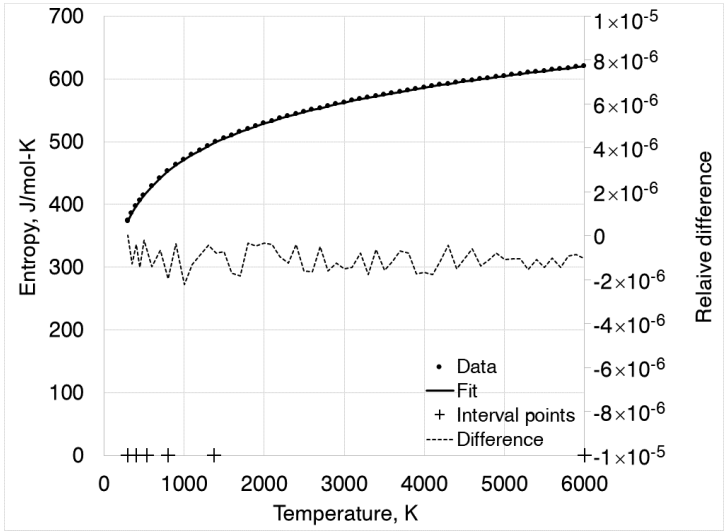


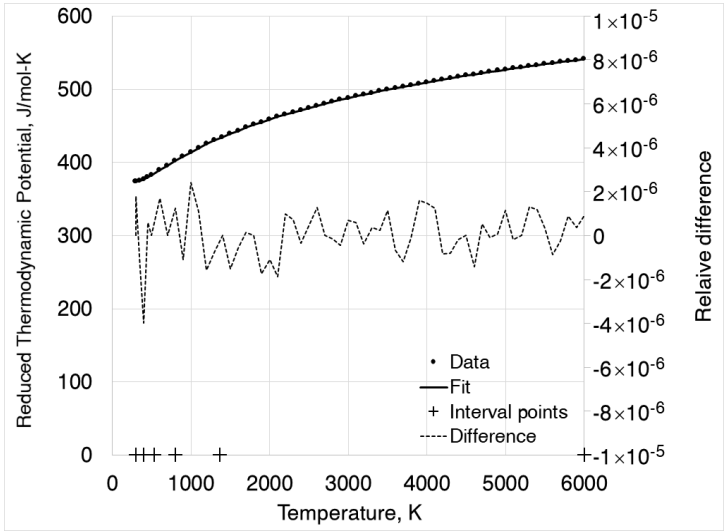
Aluminum Iodide (AlI_3) – Al1I3(g)

MW = 407.695 g/mol – $\Delta_f H^o$(298.15 K) = -193.301 kJ/mol – S^o(298.15 K) = 373.616 J/mol-K – p^o = 0.1 MPa



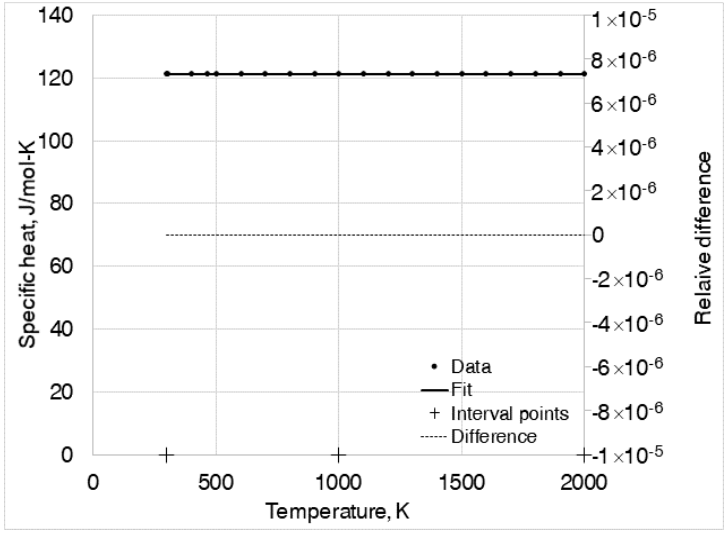


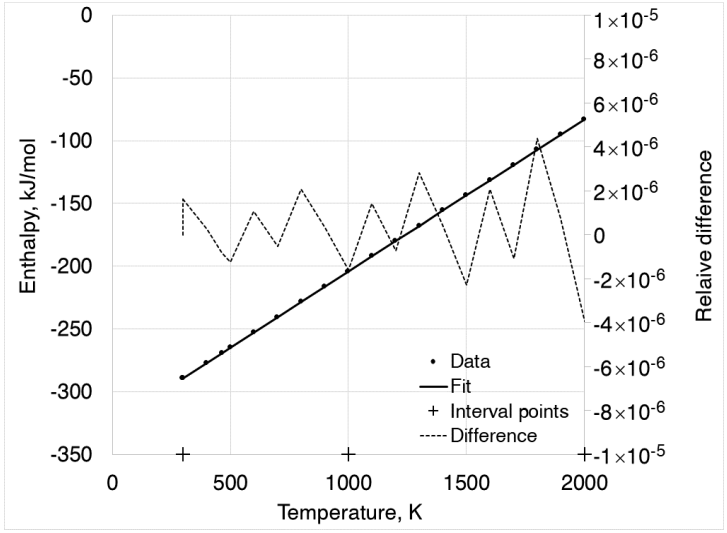


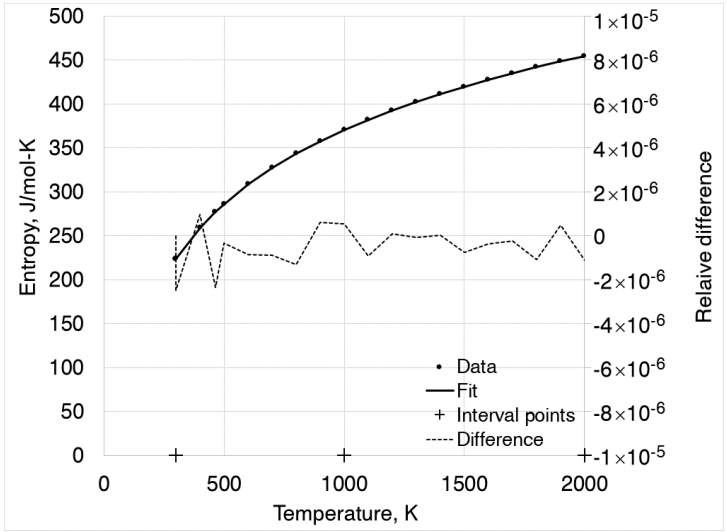


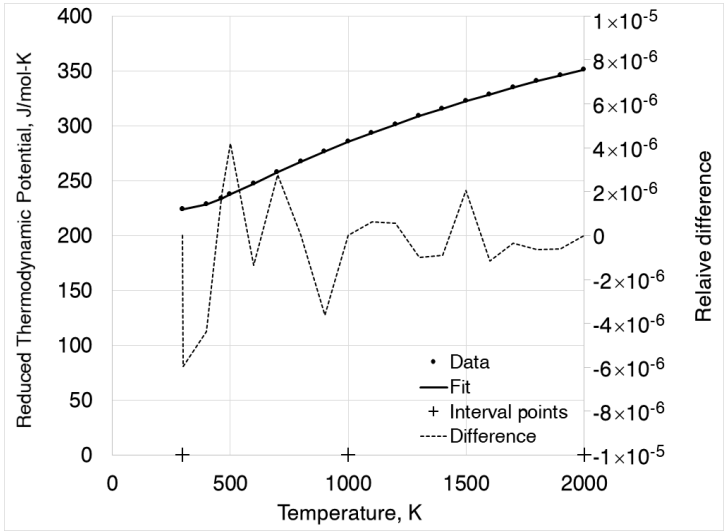
Aluminum Iodide (AlI_3) – Al1I3(l)

MW = 407.695 g/mol – $\Delta_f H^o$(298.15 K) = -289.444 kJ/mol – S^o(298.15 K) = 223.648 J/mol-K – p^o = 0.1 MPa



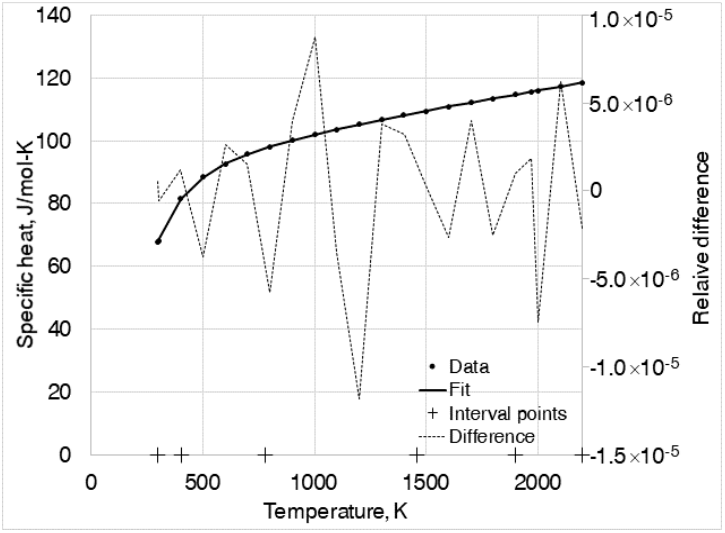


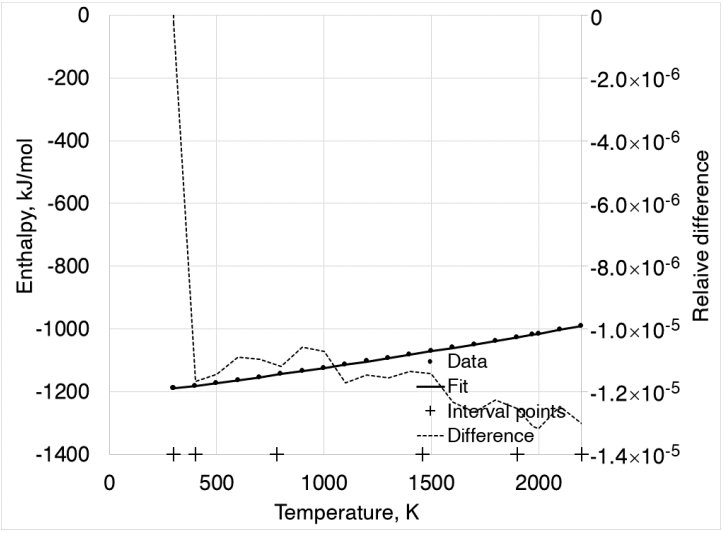


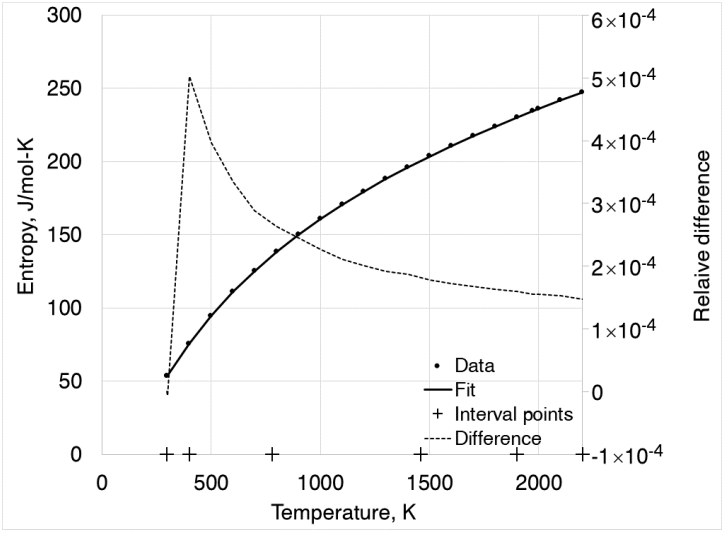


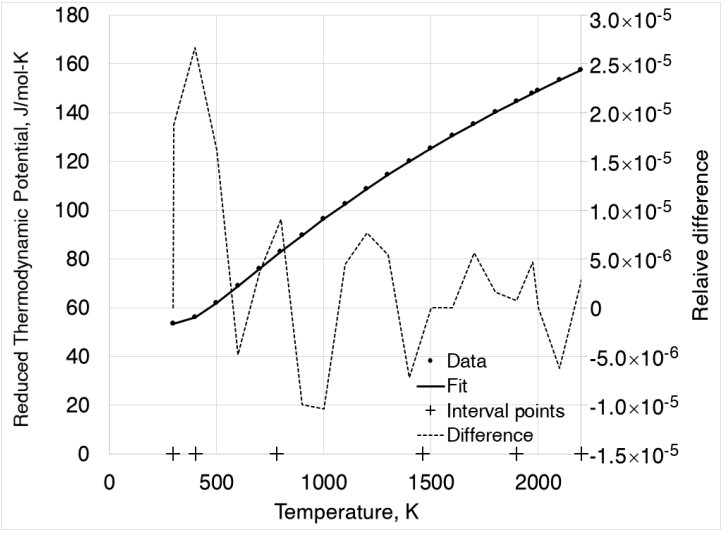
Lithium Aluminum Oxide ($LiAlO_2$) – Al1Li1O2(cr)

MW = 65.921 g/mol – $\Delta_f H^o$(298.15 K) = -1188.674 kJ/mol – S^o(298.15 K) = 53.313 J/mol-K – p^o = 0.1 MPa



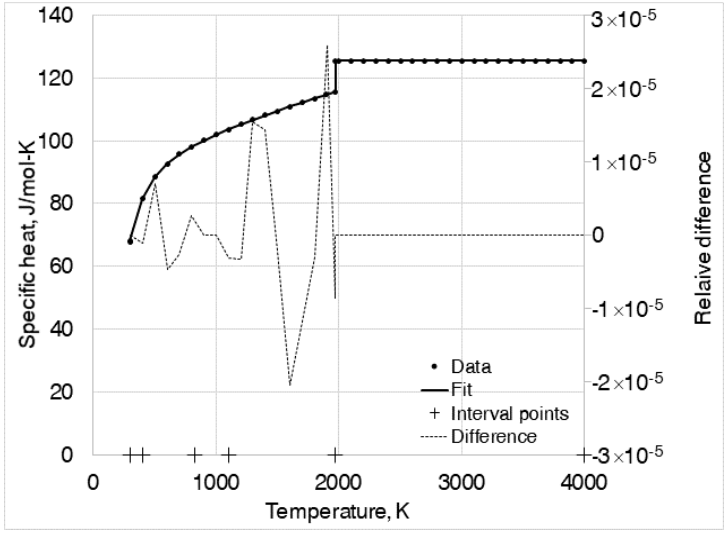


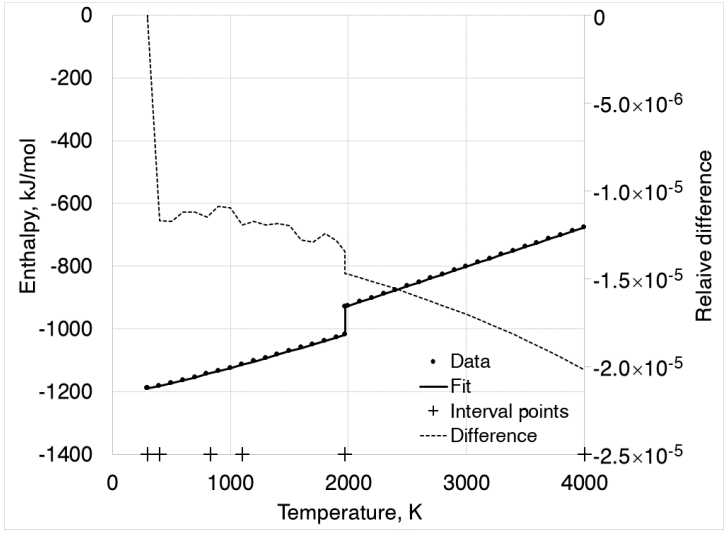


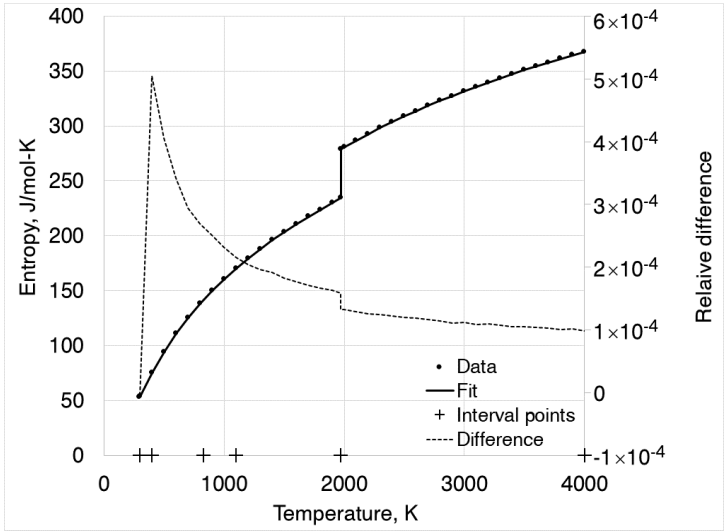


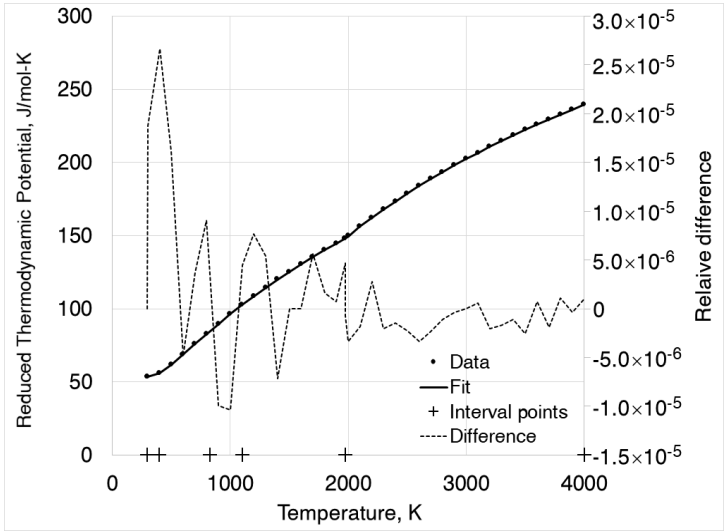
Lithium Aluminum Oxide ($LiAlO_2$) – Al1Li1O2(cr,l)

MW = 65.921 g/mol – $\Delta_f H^o$(298.15 K) = -1188.674 kJ/mol – S^o(298.15 K) = 53.313 J/mol-K – p^o = 0.1 MPa



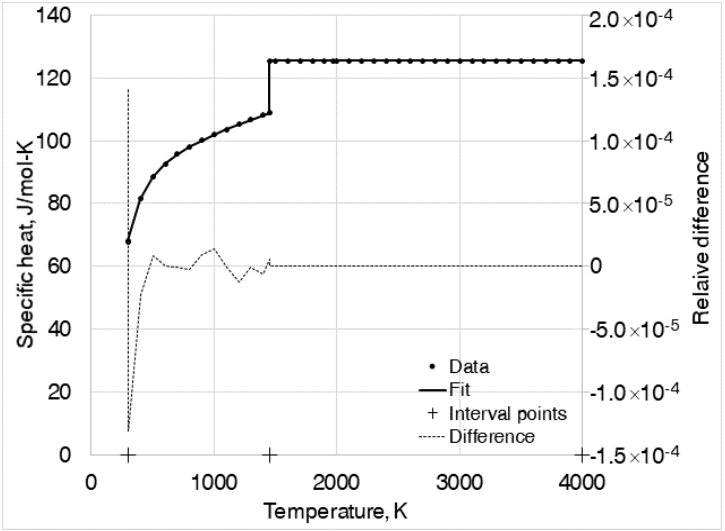


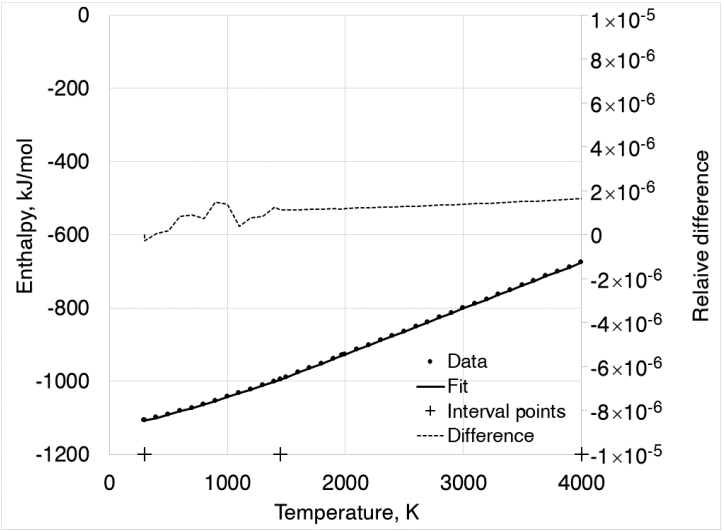


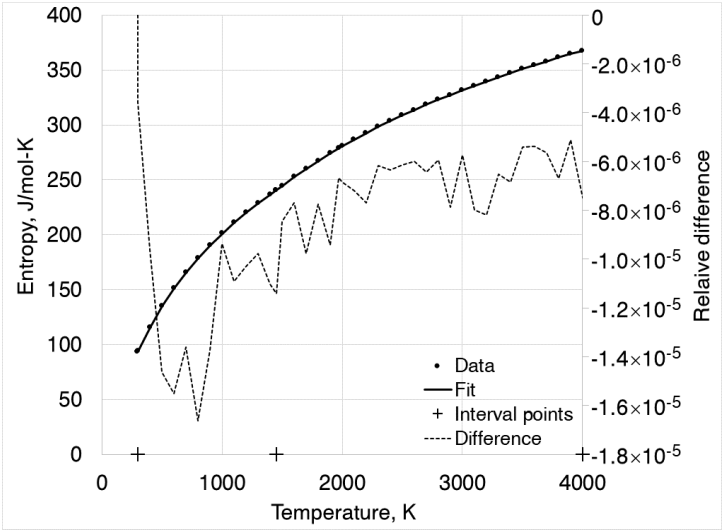


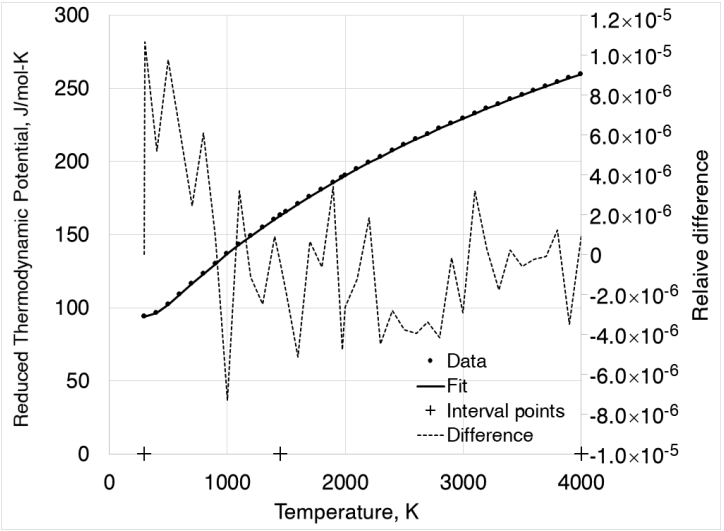
Lithium Aluminum Oxide ($LiAlO_2$) – Al1Li1O2(l)

MW = 65.921 g/mol – $\Delta_f H^o$(298.15 K) = -1107.739 kJ/mol – S^o(298.15 K) = 93.711 J/mol-K – p^o = 0.1 MPa



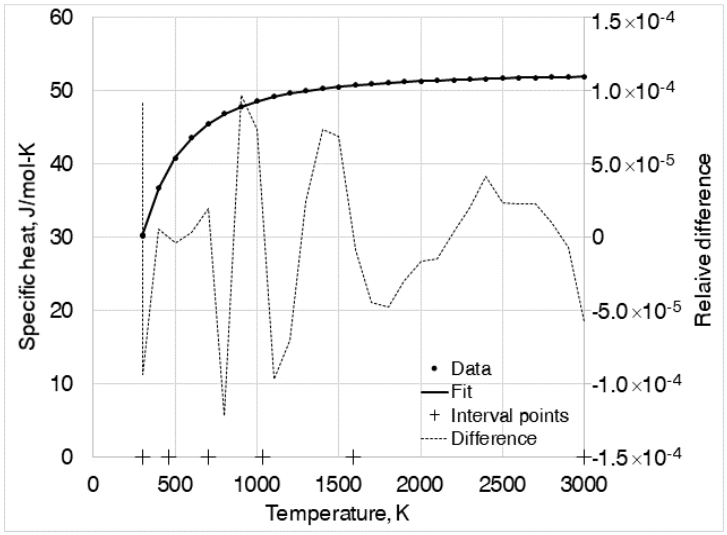


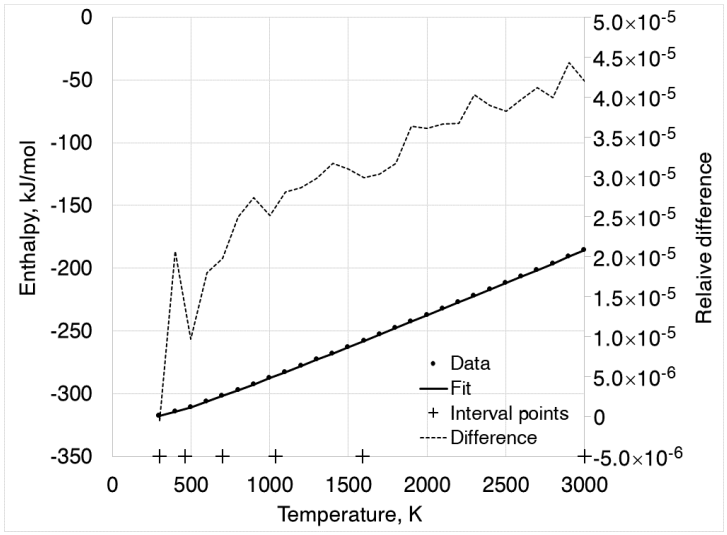


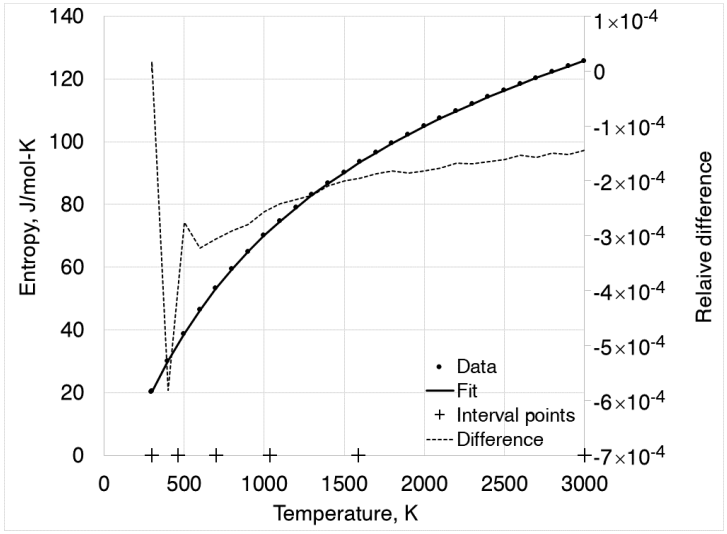


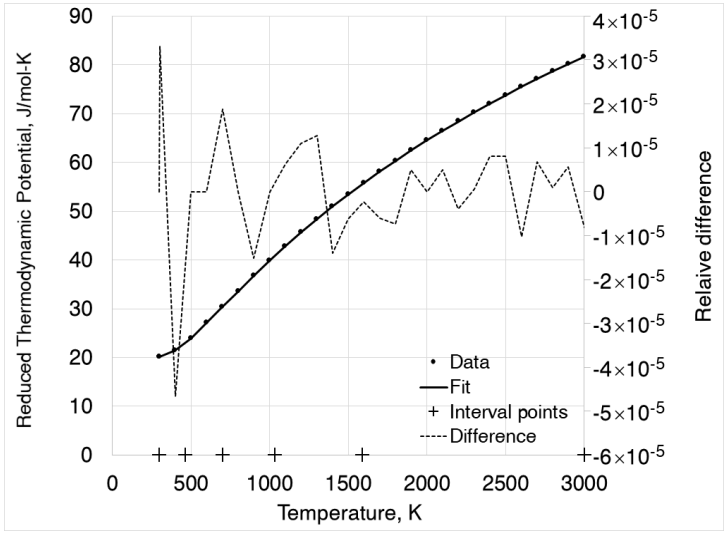
Aluminum Nitride (AlN) – Al1N1(cr)

MW = 40.9882 g/mol – $\Delta_f H^o$(298.15 K) = -317.984 kJ/mol – S^o(298.15 K) = 20.142 J/mol-K – p^o = 0.1 MPa



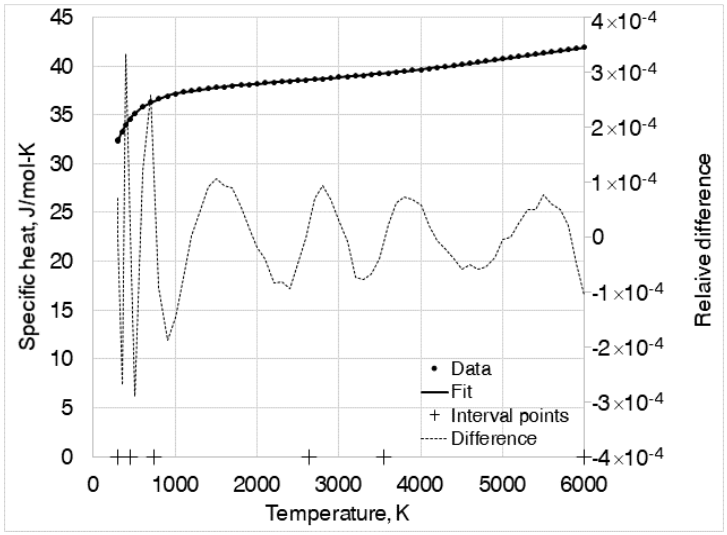


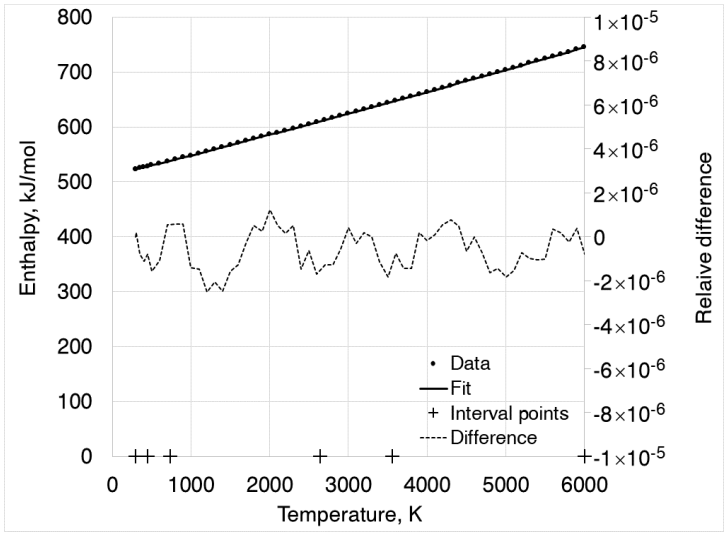


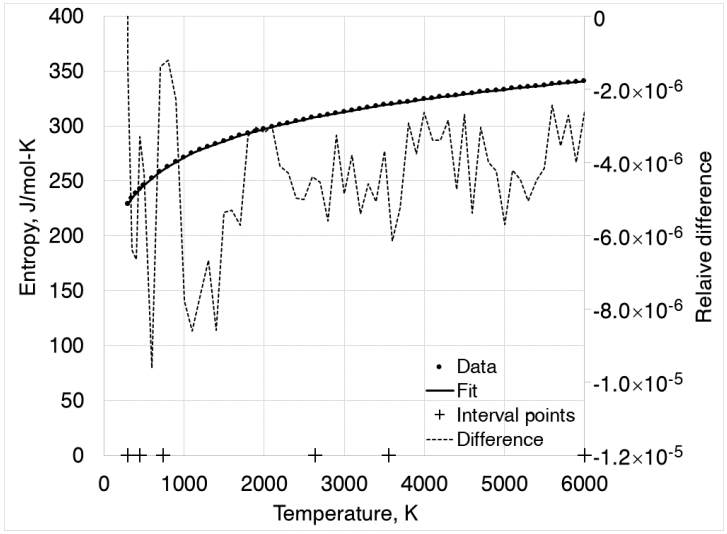


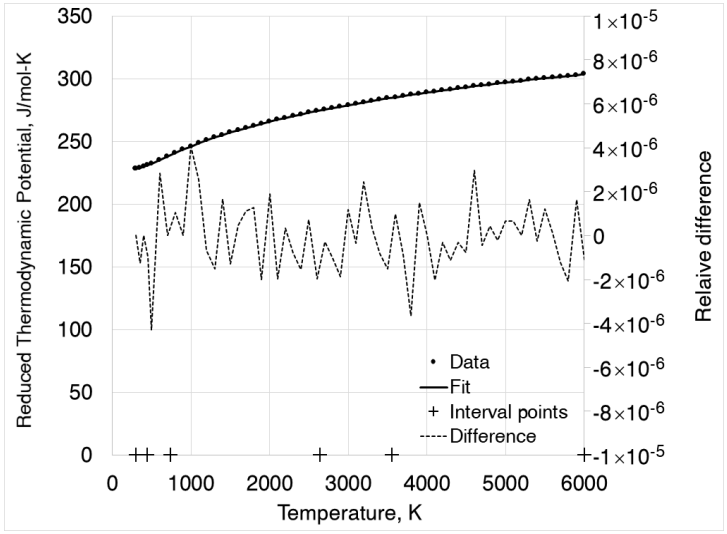
Aluminum Nitride (AlN) – Al1N1(g)

MW = 40.9882 g/mol – $\Delta_f H^o$(298.15 K) = 523 kJ/mol – S^o(298.15 K) = 228.553 J/mol-K – p^o = 0.1 MPa



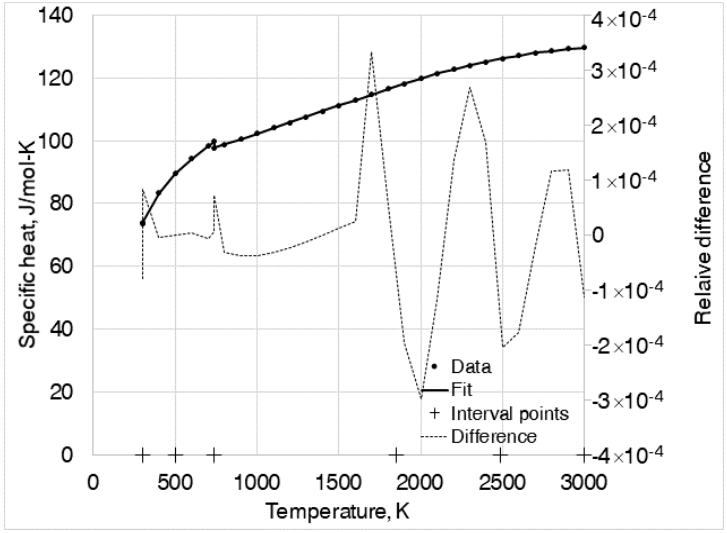


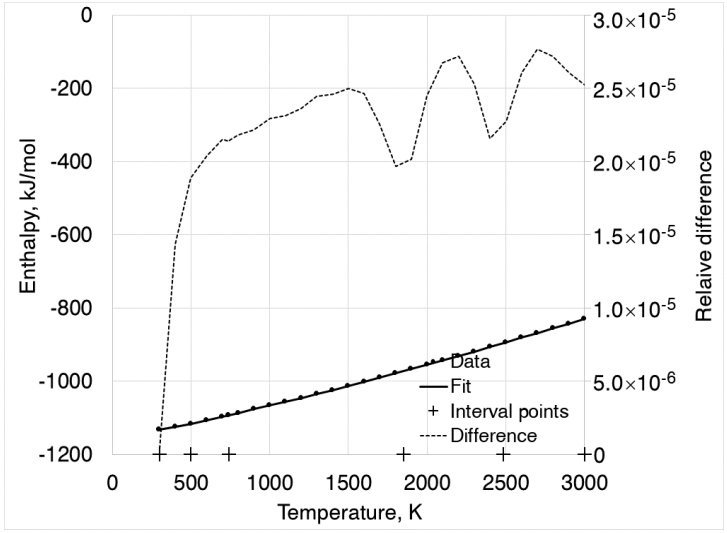


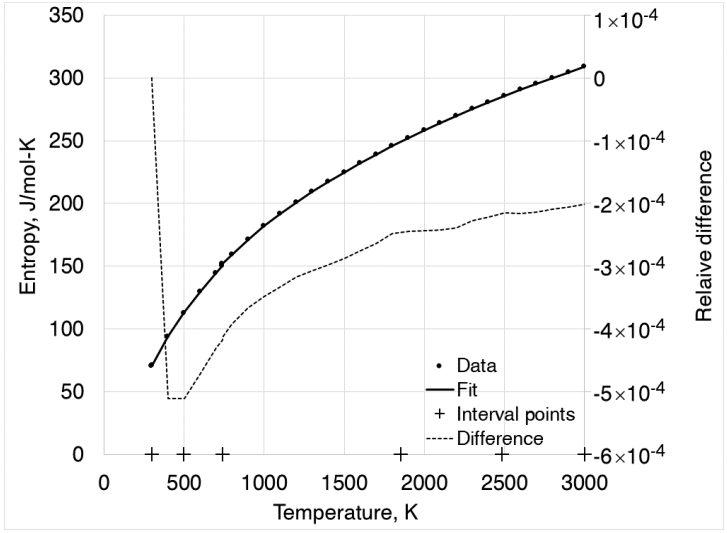


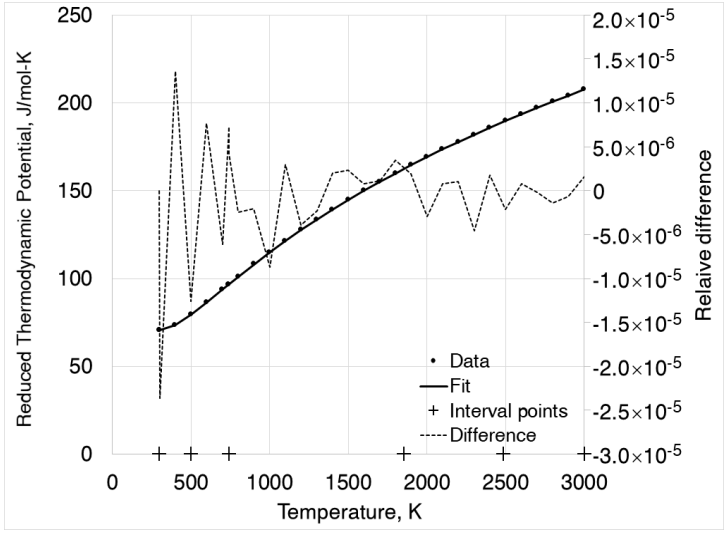
Sodium Aluminum Oxide ($NaAlO_2$) – Al1Na1O2(cr)

MW = 81.9701 g/mol – $\Delta_f H^o$(298.15 K) = -1133.195 kJ/mol – S^o(298.15 K) = 70.4 J/mol-K – p^o = 0.1 MPa



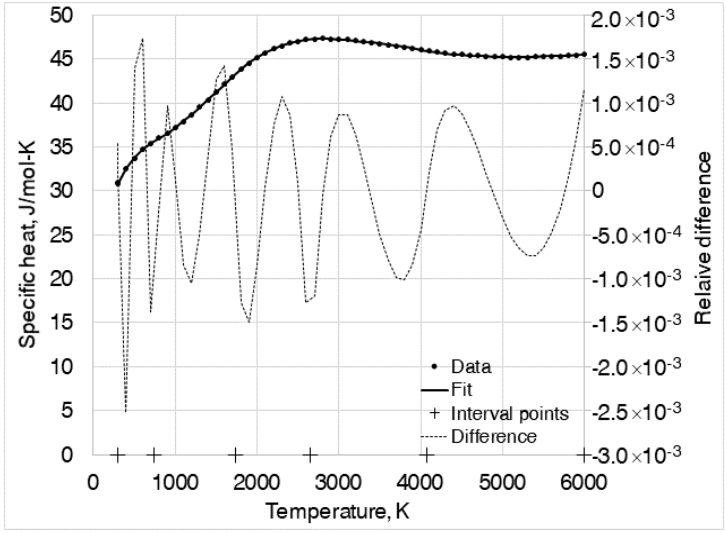


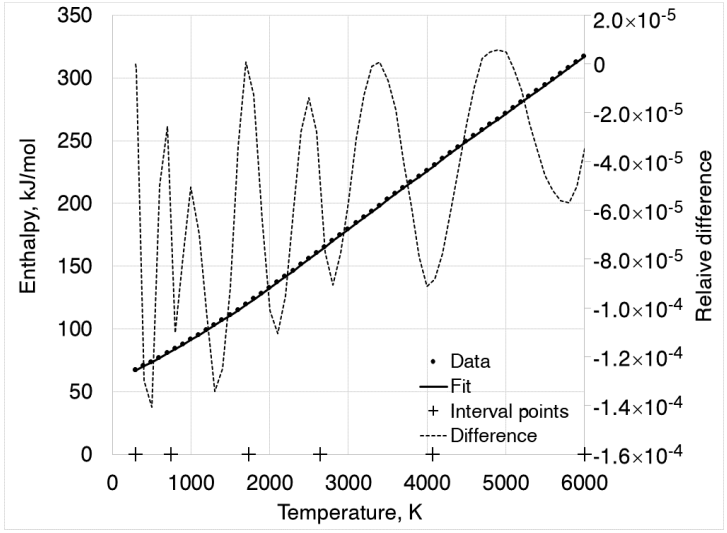


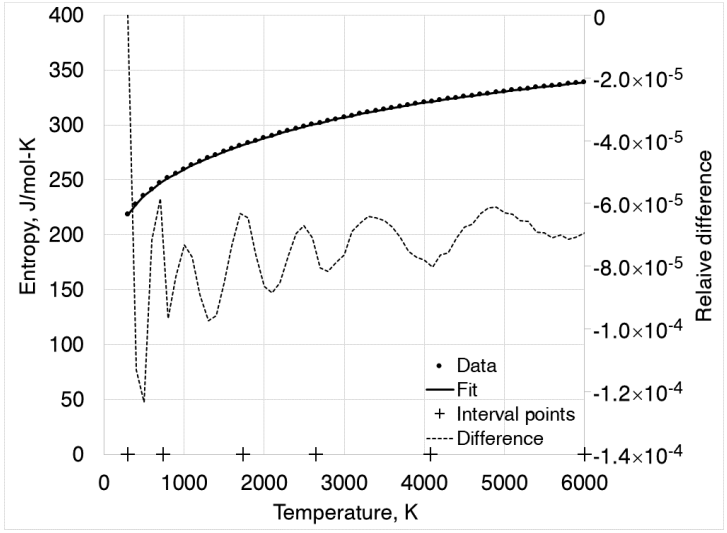


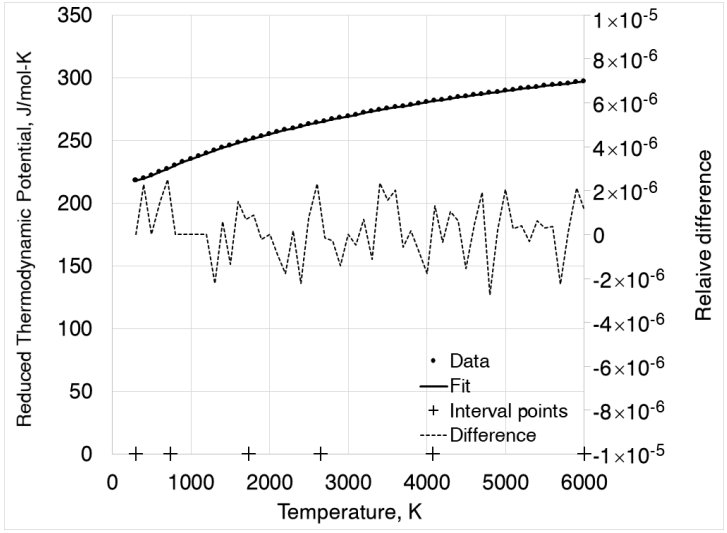
Aluminum Oxide (AlO) – Al1O1(g)

MW = 42.9809 g/mol – $\Delta_f H^o$(298.15 K) = 66.944 kJ/mol – S^o(298.15 K) = 218.386 J/mol-K – p^o = 0.1 MPa



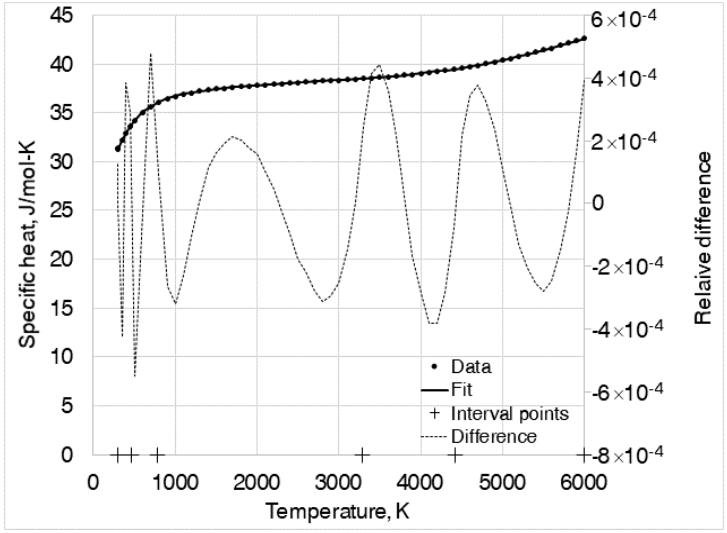


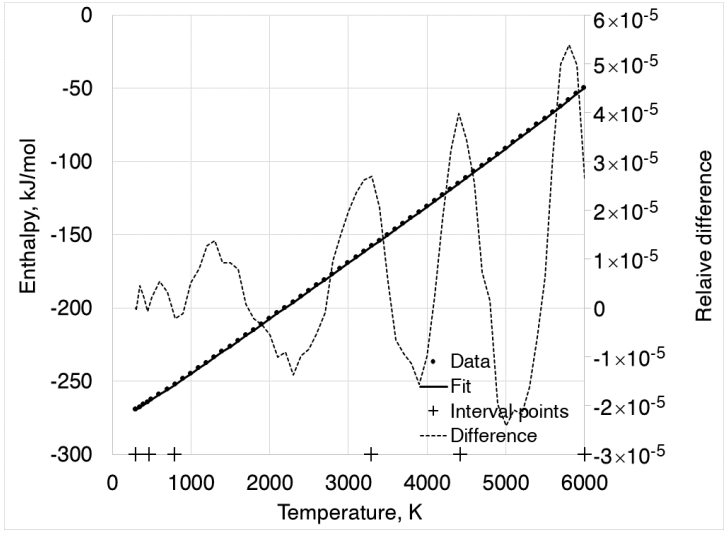


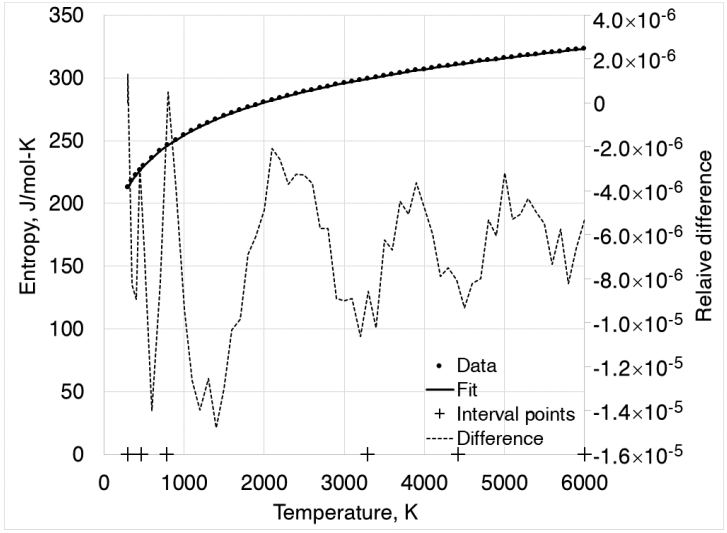


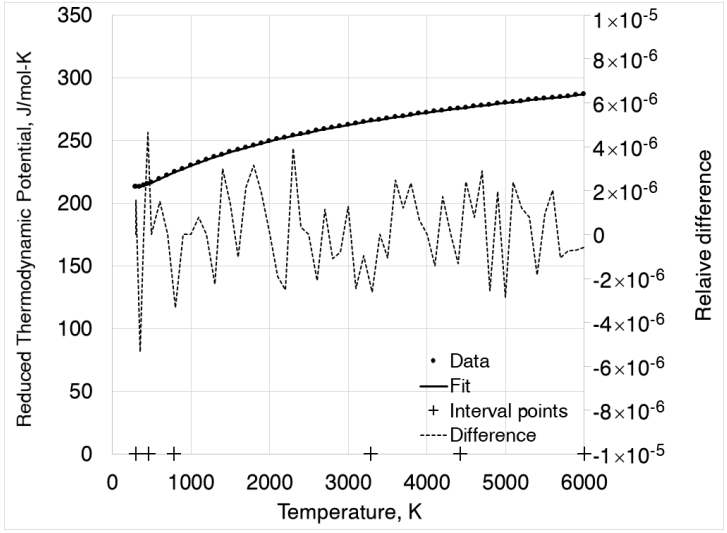
Aluminum Oxide, Ion (AlO^-) – Al1O1-(g)

MW = 42.9815 g/mol – $\Delta_f H^o$(298.15 K) = -269.424 kJ/mol – S^o(298.15 K) = 212.807 J/mol-K – p^o = 0.1 MPa



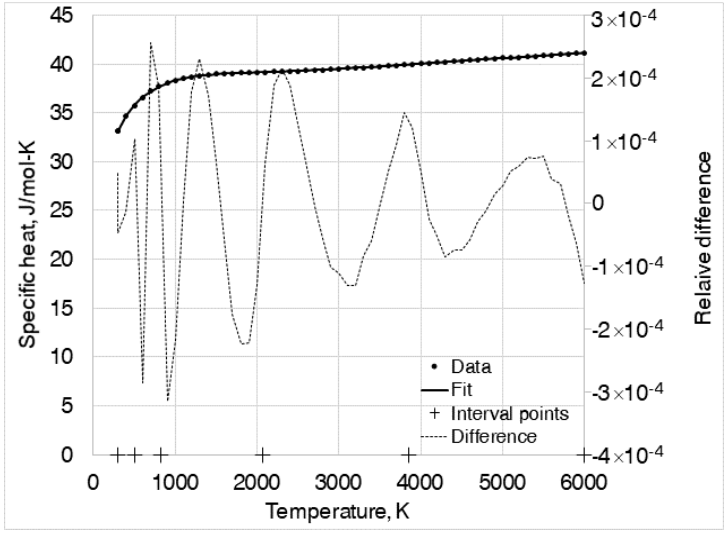


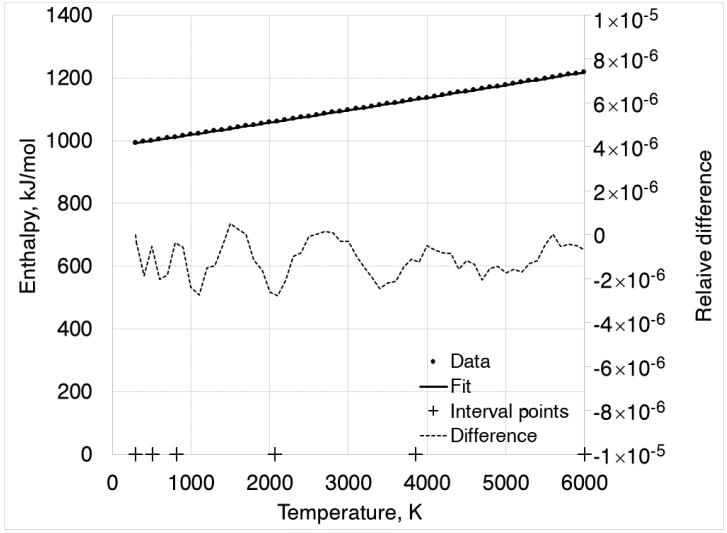


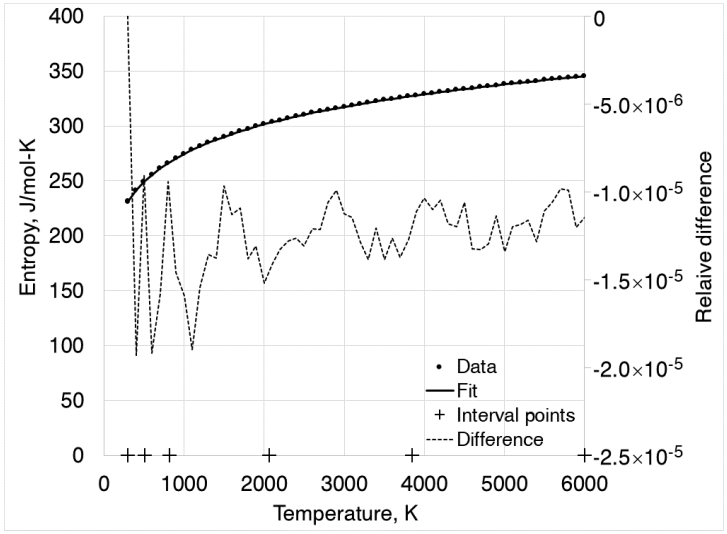


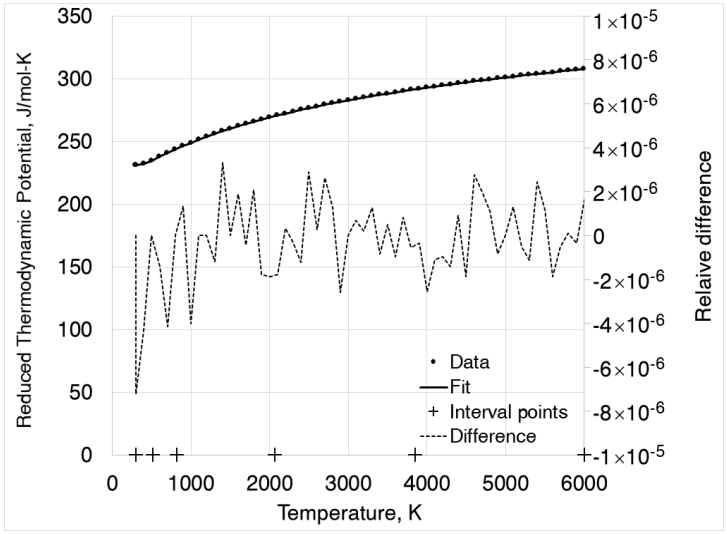
Aluminum Oxide, Ion (AlO^+) – Al1O1+(g)

MW = 42.9804 g/mol – $\Delta_f H^o$(298.15 K) = 992.993 kJ/mol – S^o(298.15 K) = 230.975 J/mol-K – p^o = 0.1 MPa



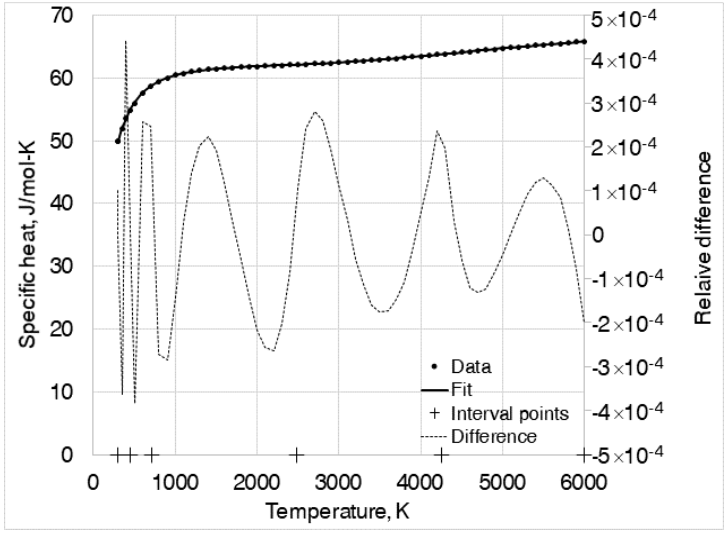


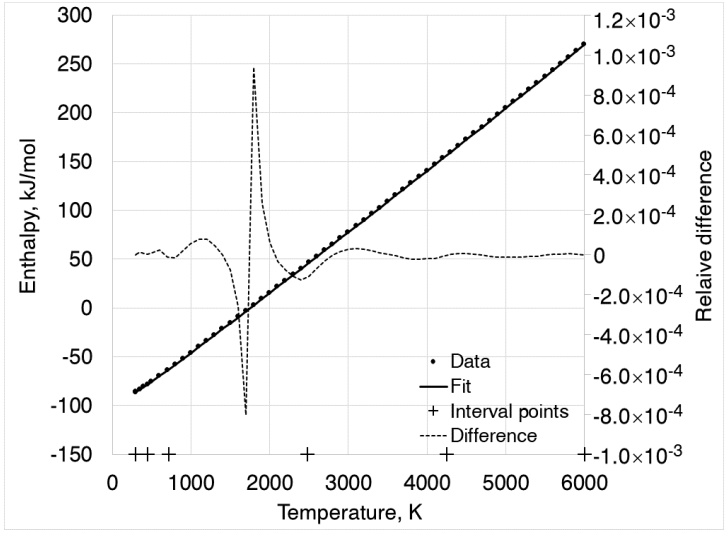


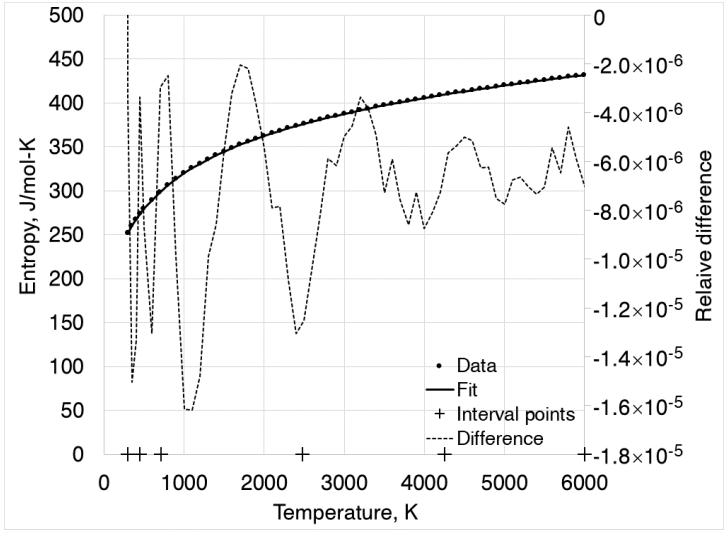


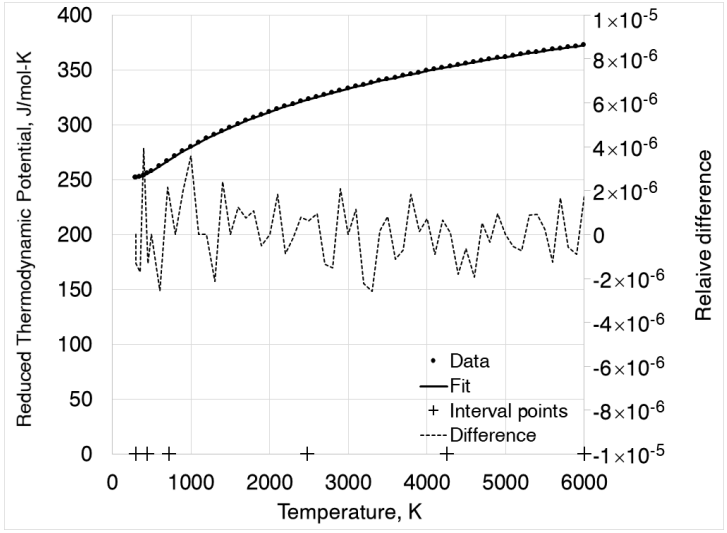
Aluminum Oxide (AlO_2) – Al1O2(g)

MW = 58.9803 g/mol – $\Delta_f H^o$(298.15 K) = -86.19 kJ/mol – S^o(298.15 K) = 251.832 J/mol-K – p^o = 0.1 MPa



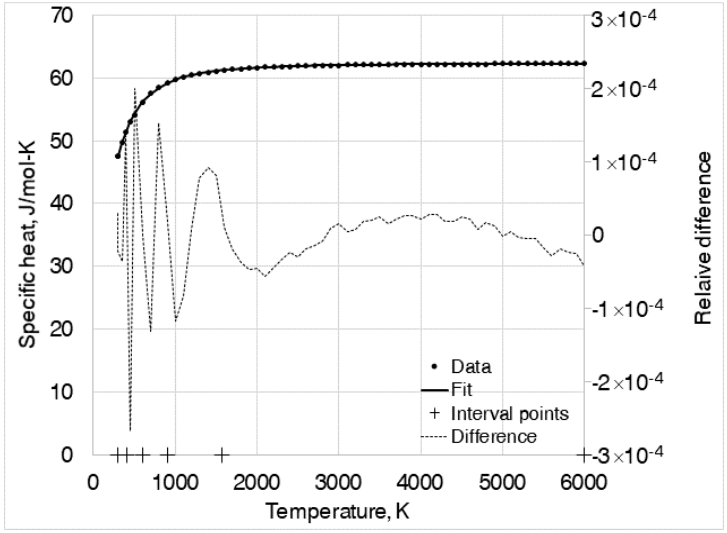


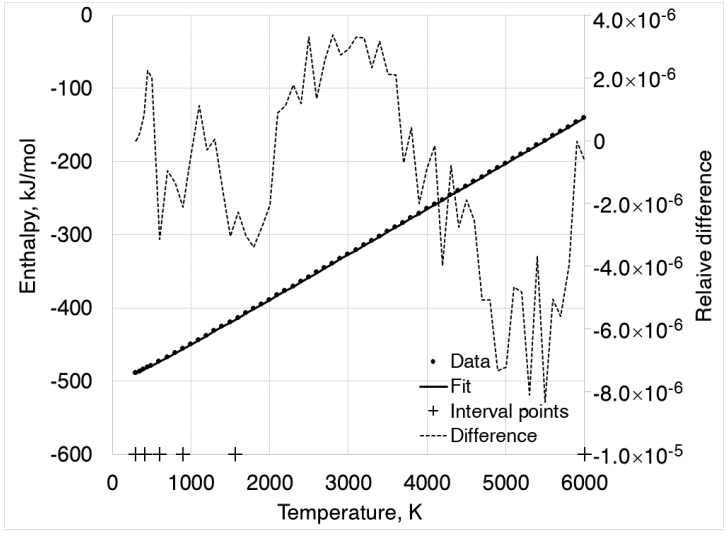


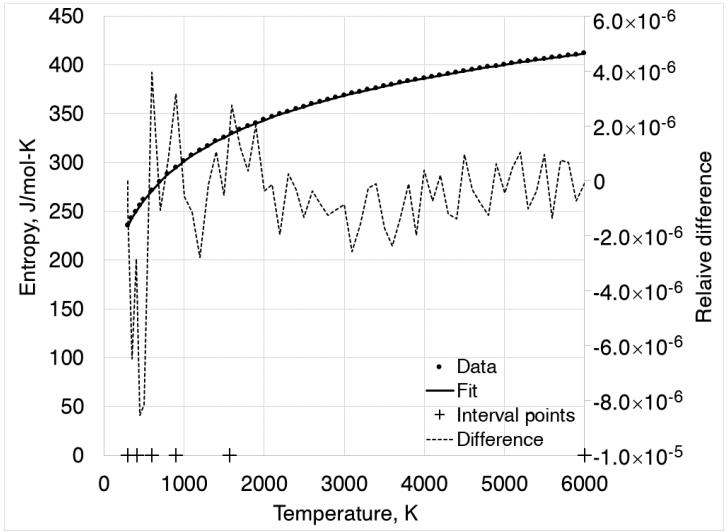


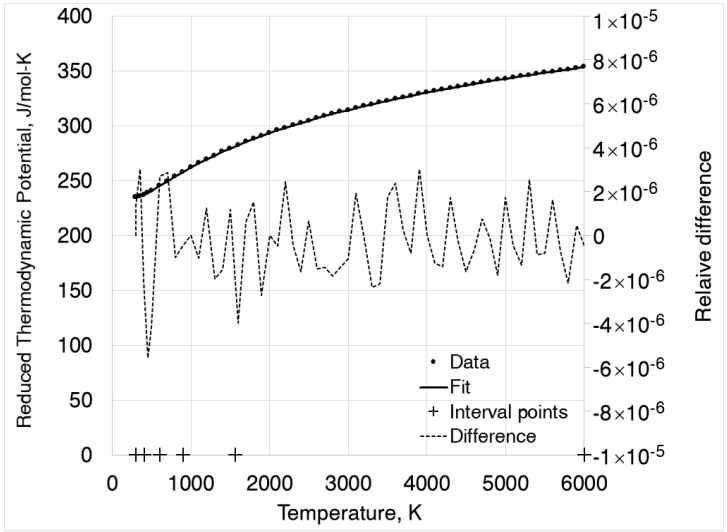
Aluminum Oxide, Ion (AlO_2^-) – Al1O2-(g)

MW = 58.9809 g/mol – $\Delta_f H^o$(298.15 K) = -489.21 kJ/mol – S^o(298.15 K) = 235.094 J/mol-K – p^o = 0.1 MPa



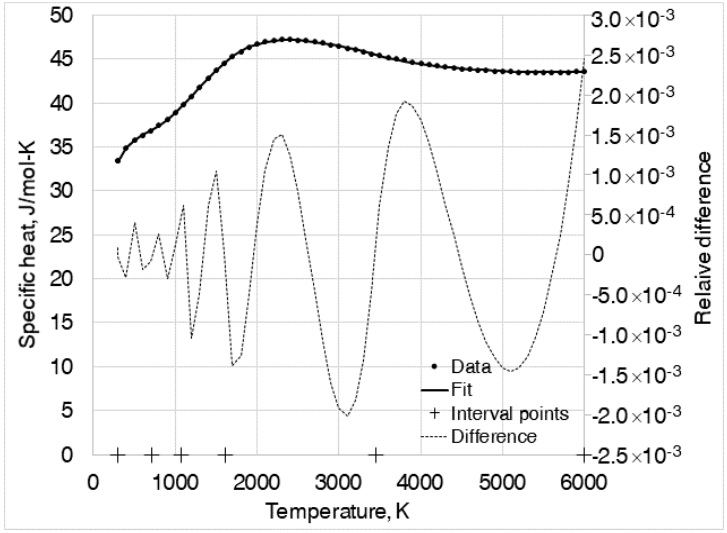


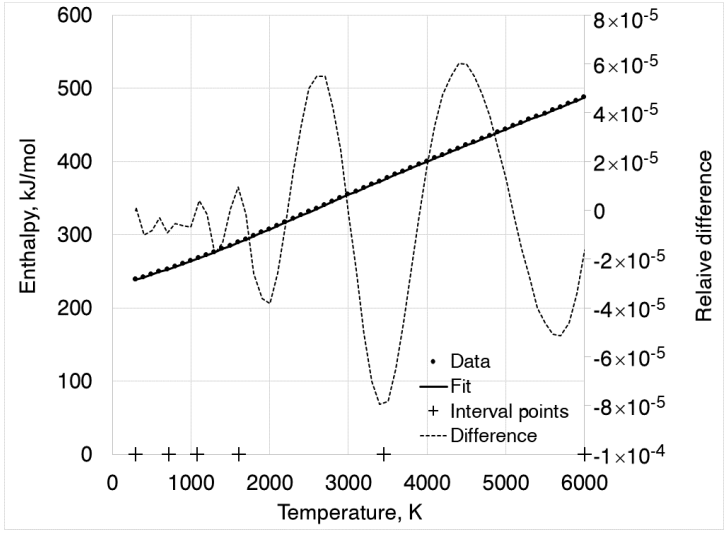


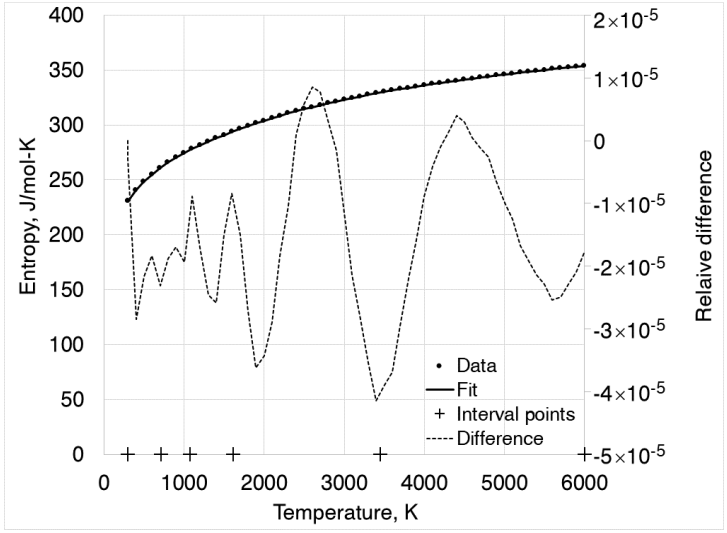


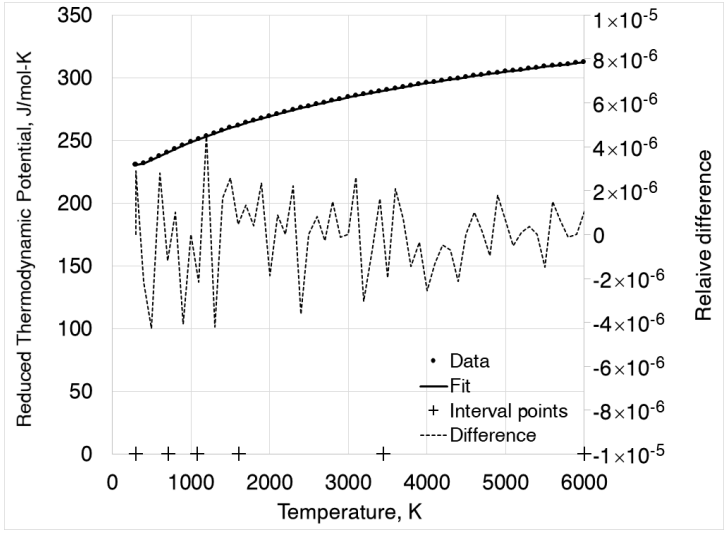
Aluminum Sulfide (AlS) – Al1S1(g)

MW = 59.04 g/mol – $\Delta_f H^o$(298.15 K) = 238.488 kJ/mol – S^o(298.15 K) = 230.601 J/mol-K – p^o = 0.1 MPa



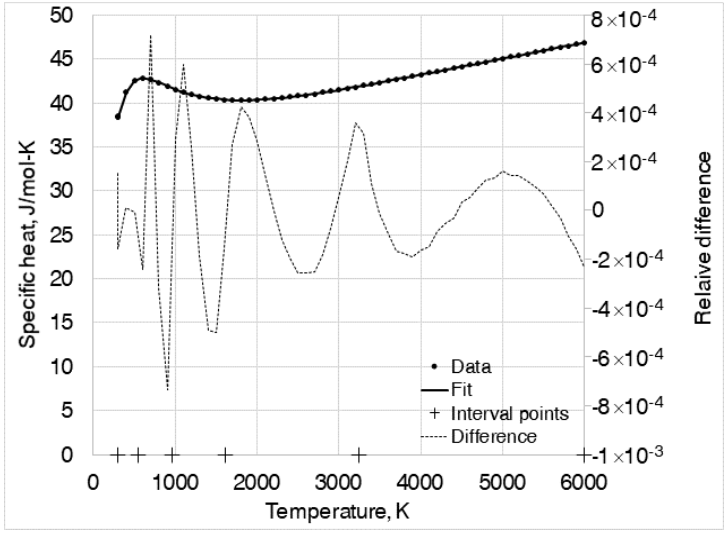


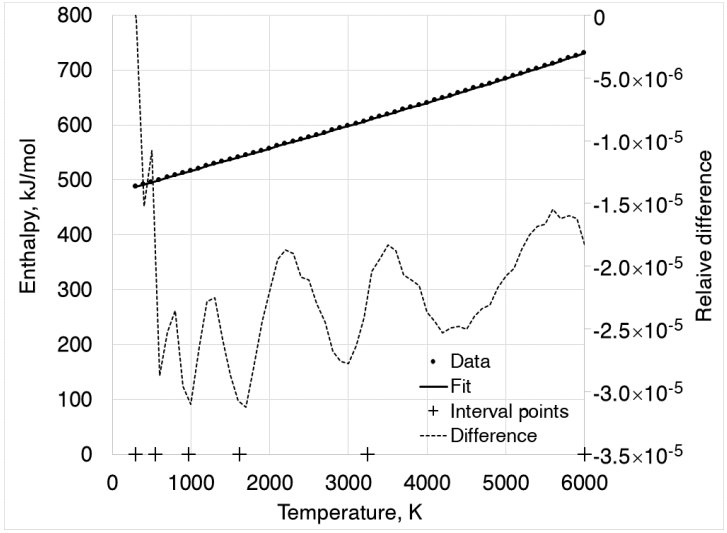


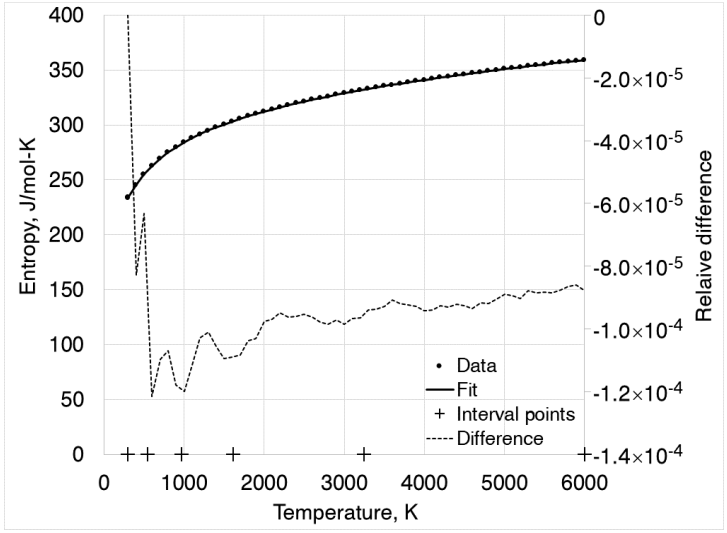


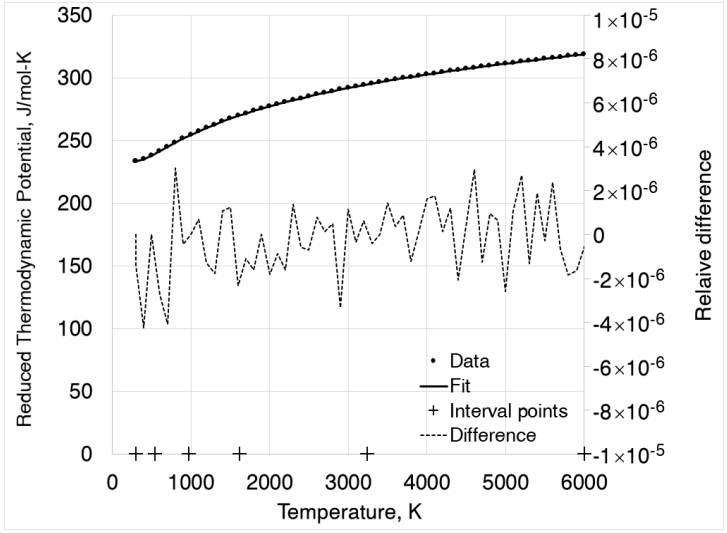
Aluminum (Al_2) – Al2(g)

MW = 53.96308 g/mol – $\Delta_f H^o$(298.15 K) = 487.018 kJ/mol – S^o(298.15 K) = 233.531 J/mol-K – p^o = 0.1 MPa



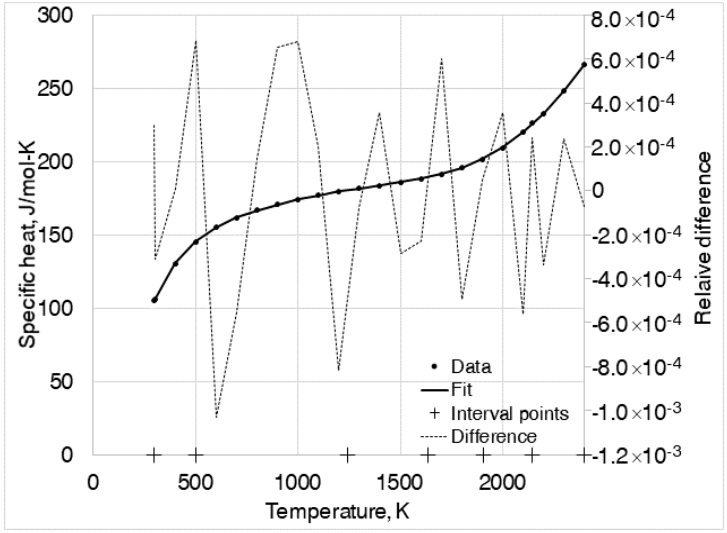


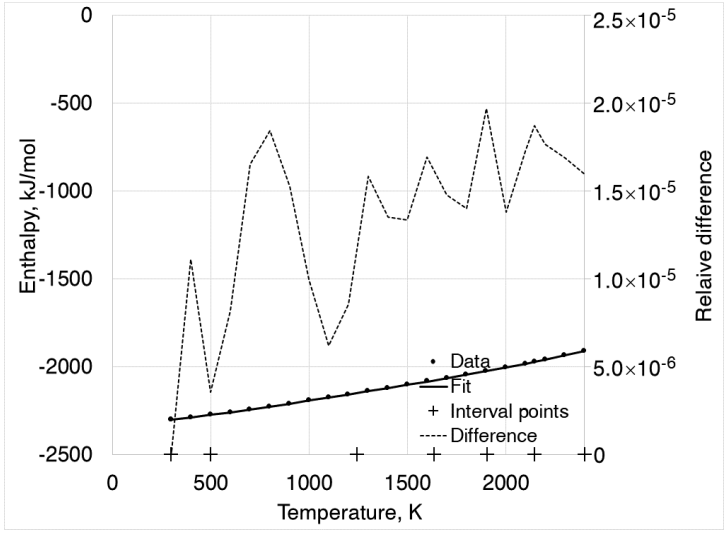


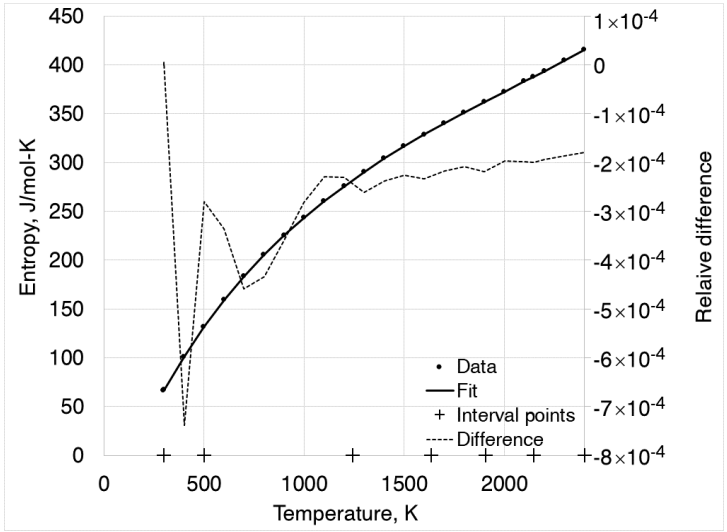


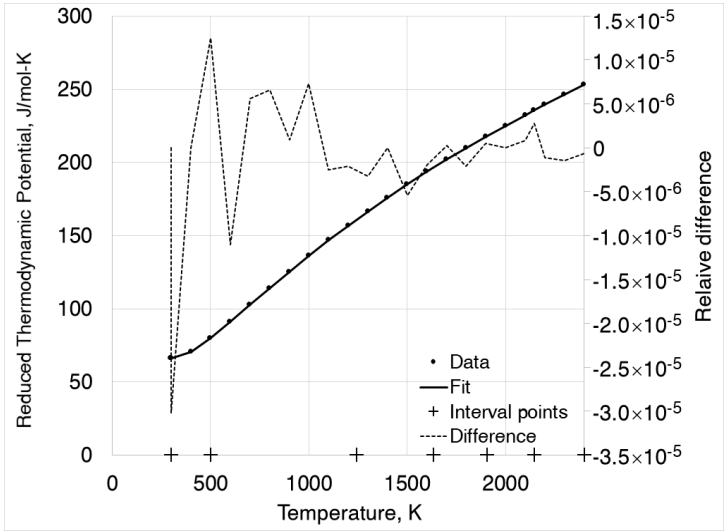
Beryllium Aluminum Oxide ($BeAl_2O_4$) – Al2Be1O4(cr)

MW = 126.9729 g/mol – $\Delta_f H^o$(298.15 K) = -2300.782 kJ/mol – S^o(298.15 K) = 66.291 J/mol-K – p^o = 0.1 MPa



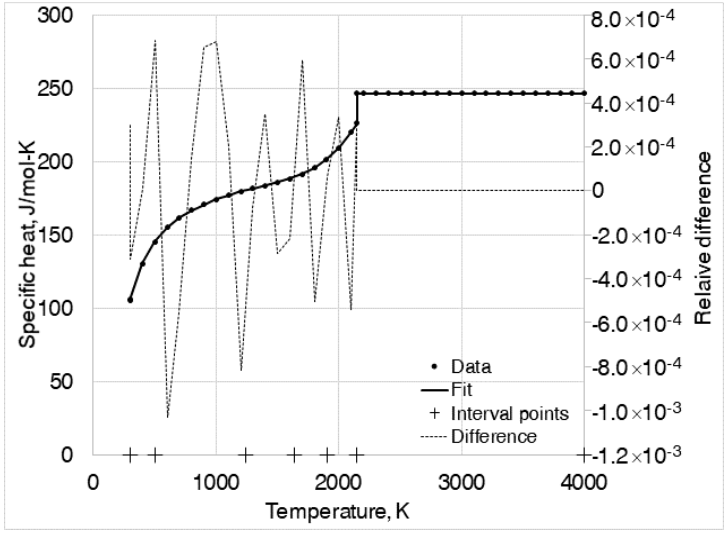


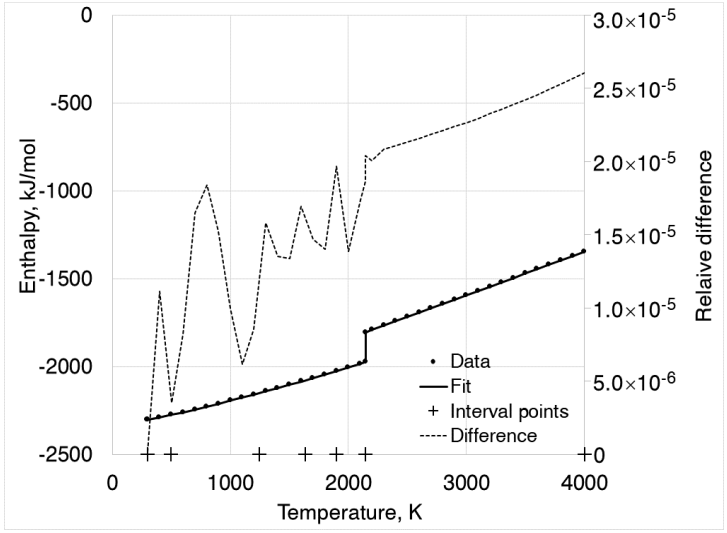


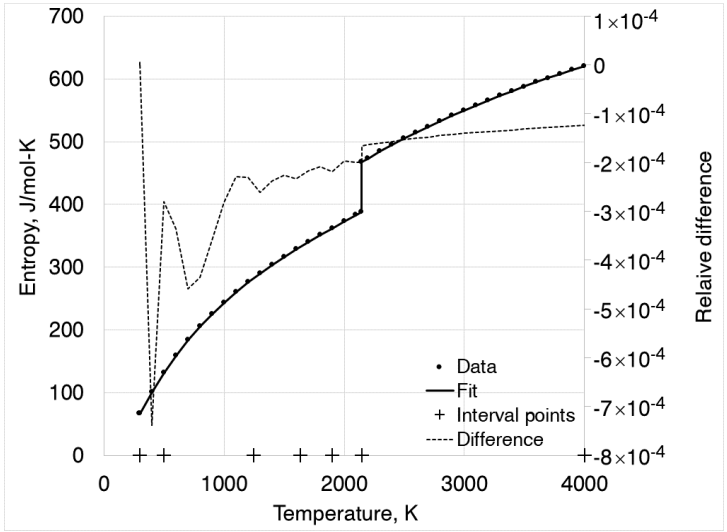


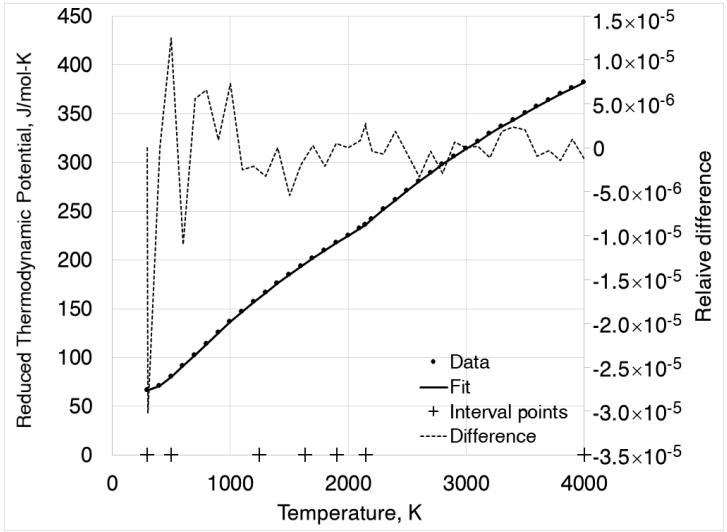
Beryllium Aluminum Oxide ($BeAl_2O_4$) – Al2Be1O4(cr,l)

MW = 126.9729 g/mol – $\Delta_f H^o$(298.15 K) = -2300.782 kJ/mol – S^o(298.15 K) = 66.291 J/mol-K – p^o = 0.1 MPa



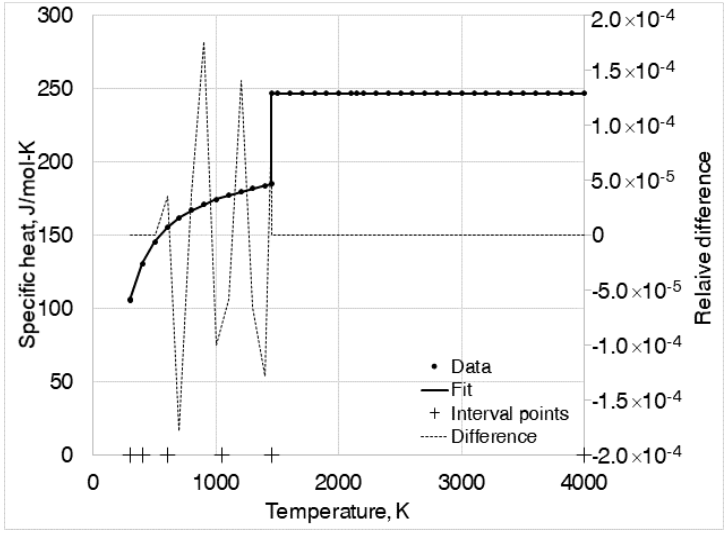


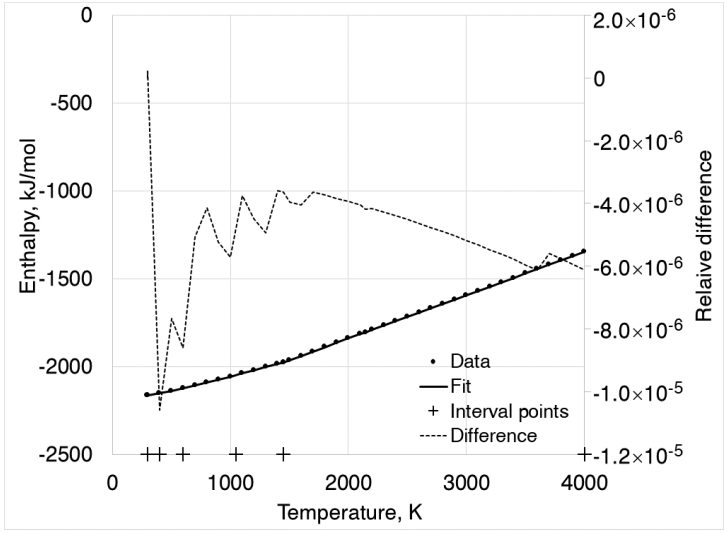


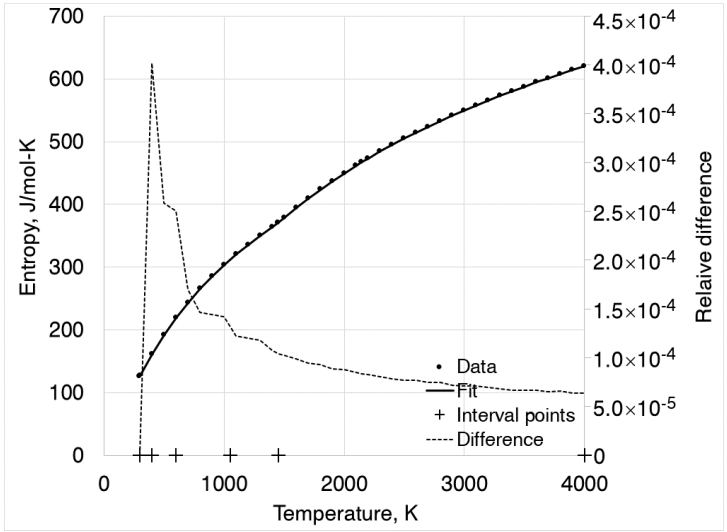


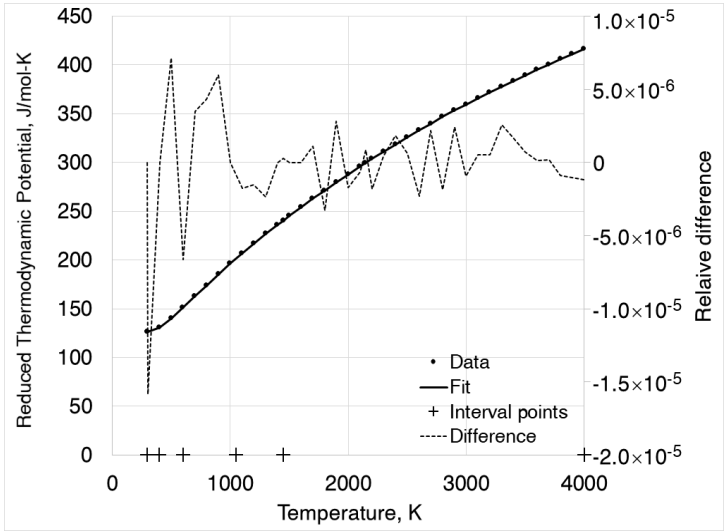
Beryllium Aluminum Oxide ($BeAl_2O_4$) – Al2Be1O4(l)

MW = 126.9729 g/mol – $\Delta_f H^o$(298.15 K) = -2163.842 kJ/mol – S^o(298.15 K) = 126.406 J/mol-K – p^o = 0.1 MPa



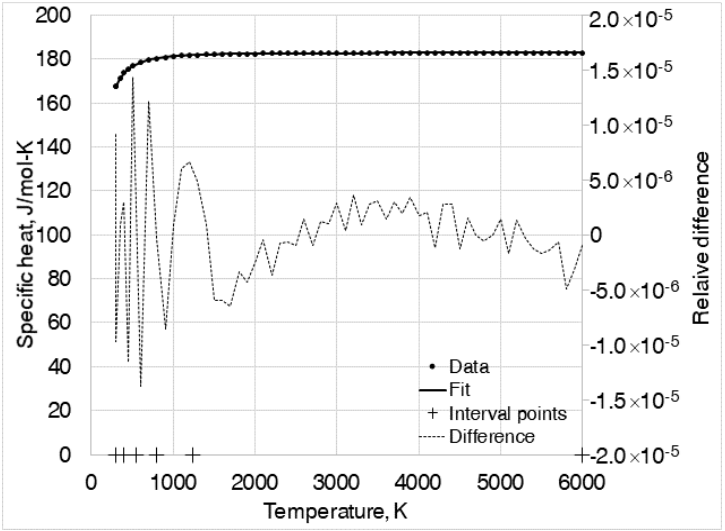


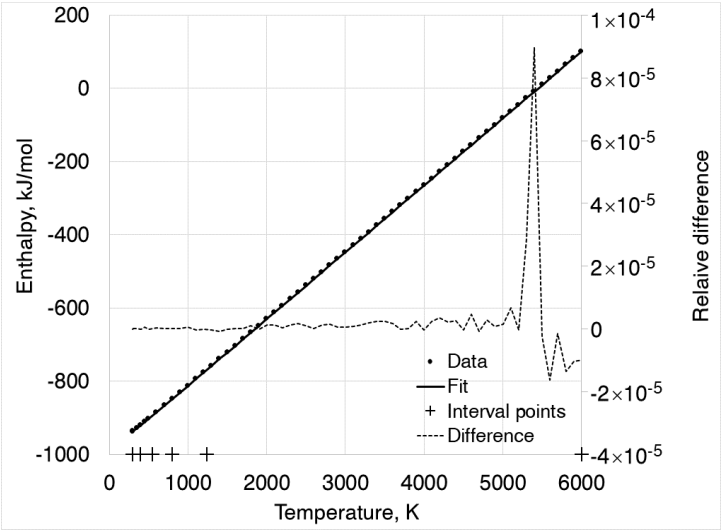


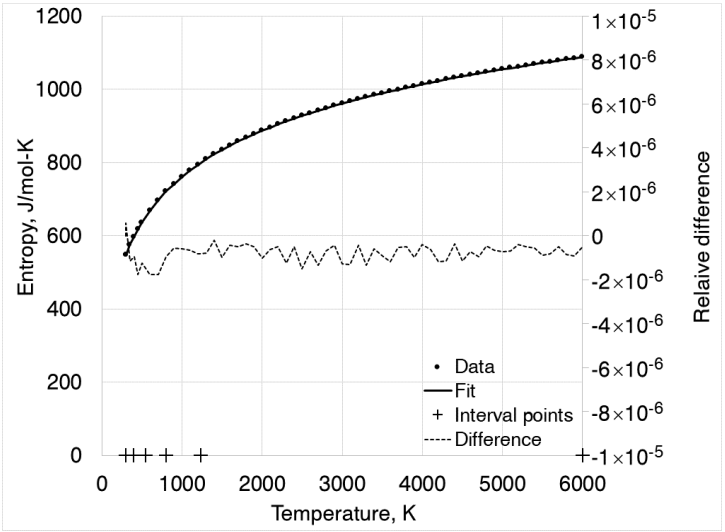


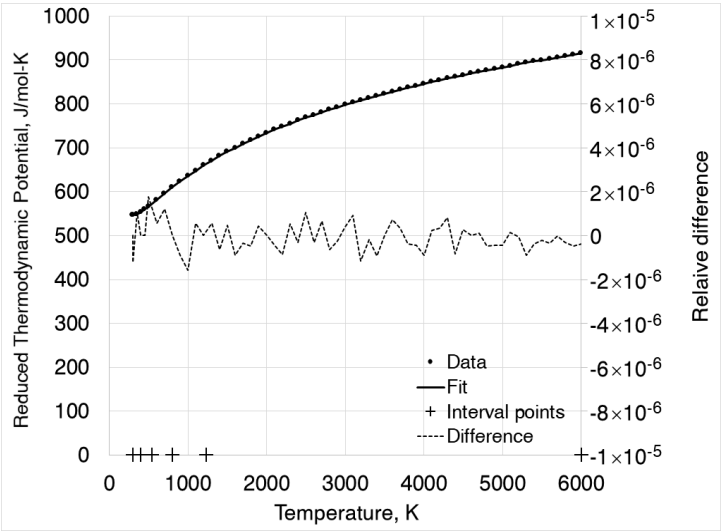
Aluminum Bromide ($(AlBr_3)_2$) – Al2Br6(g)

MW = 533.387 g/mol – $\Delta_f H^o$(298.15 K) = -937.216 kJ/mol – S^o(298.15 K) = 547.22 J/mol-K – p^o = 0.1 MPa



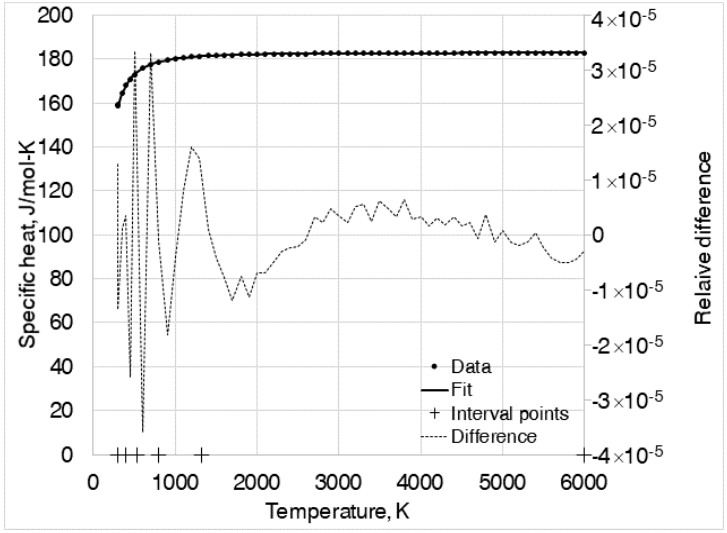


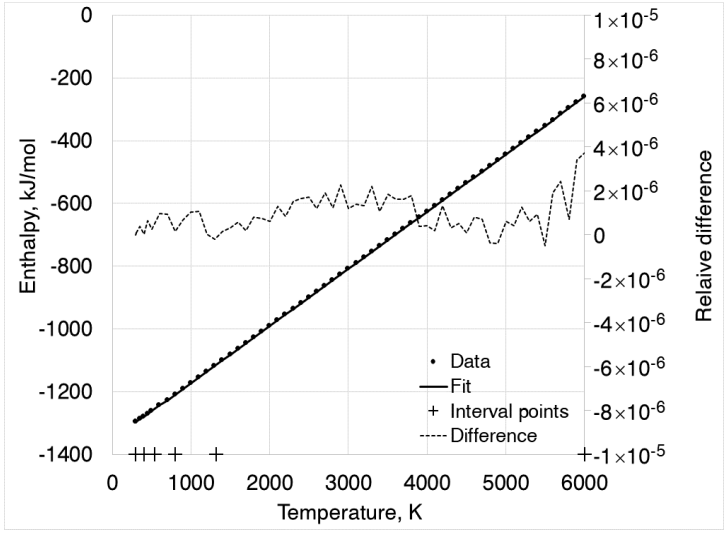


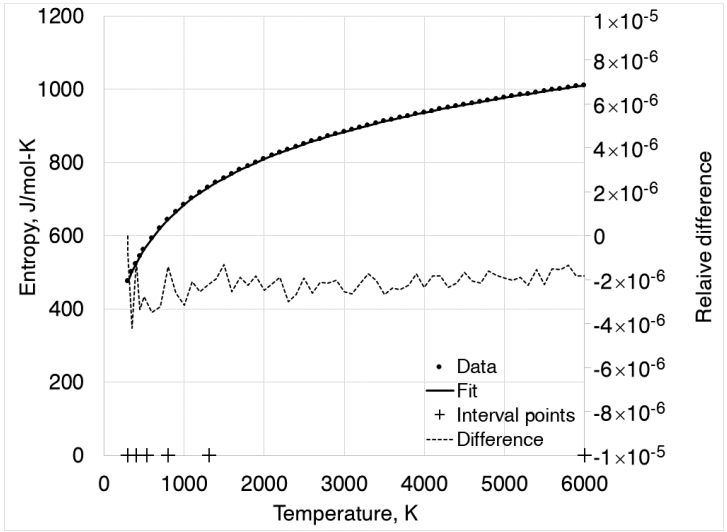


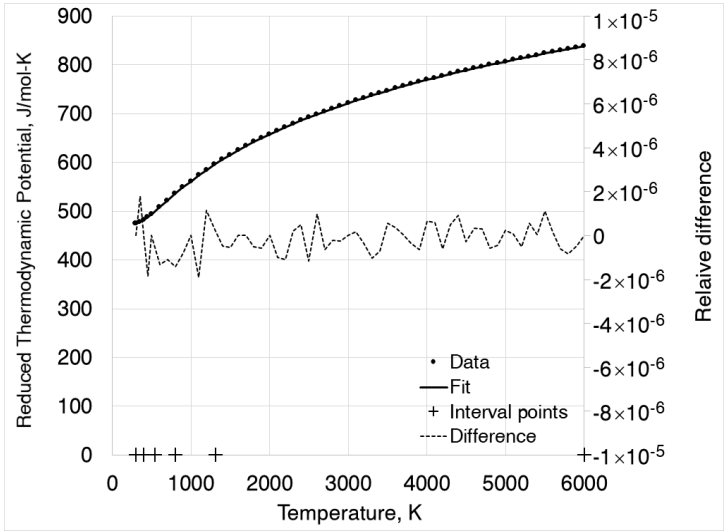
Aluminum Chloride ($(AlCl_3)_2$) – Al2Cl6(g)

MW = 266.681 g/mol – $\Delta_f H^o$(298.15 K) = -1295.743 kJ/mol – S^o(298.15 K) = 475.05 J/mol-K – p^o = 0.1 MPa



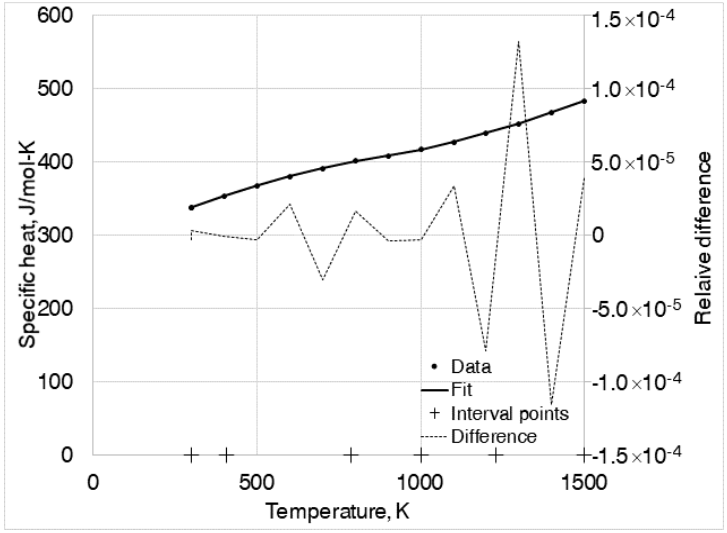


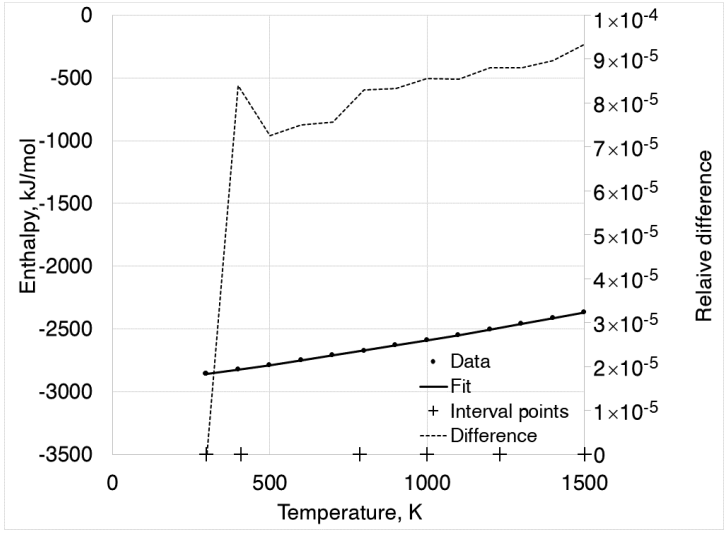


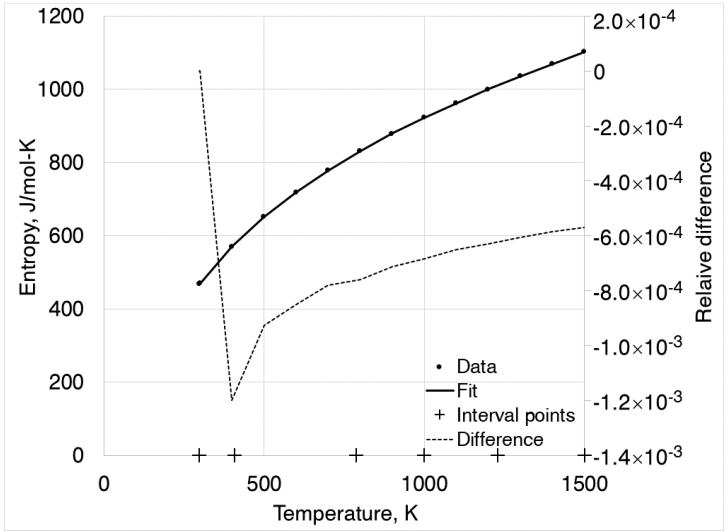


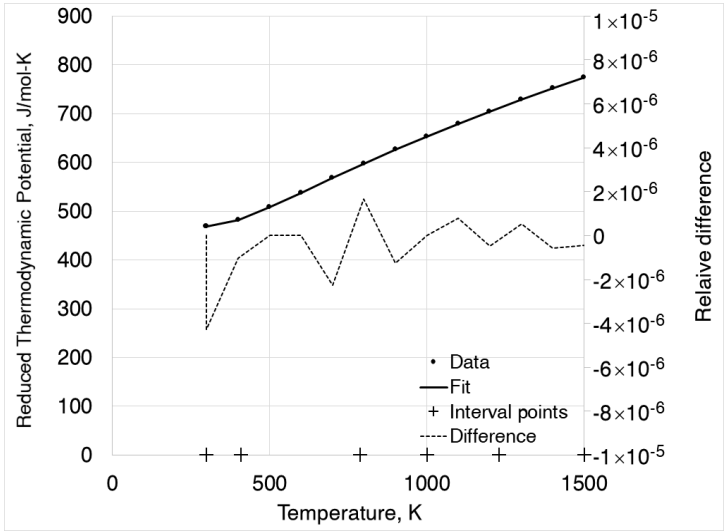
Potassium Aluminum Chloride ($K_3Al_2Cl_9$) – Al2Cl9K3(cr)

MW = 490.335 g/mol – $\Delta_f H^o$(298.15 K) = -2860.182 kJ/mol – S^o(298.15 K) = 468.608 J/mol-K – p^o = 0.1 MPa



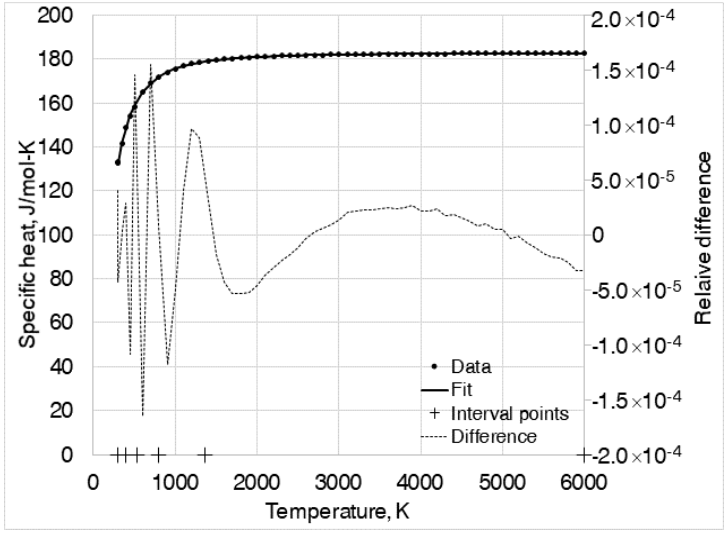


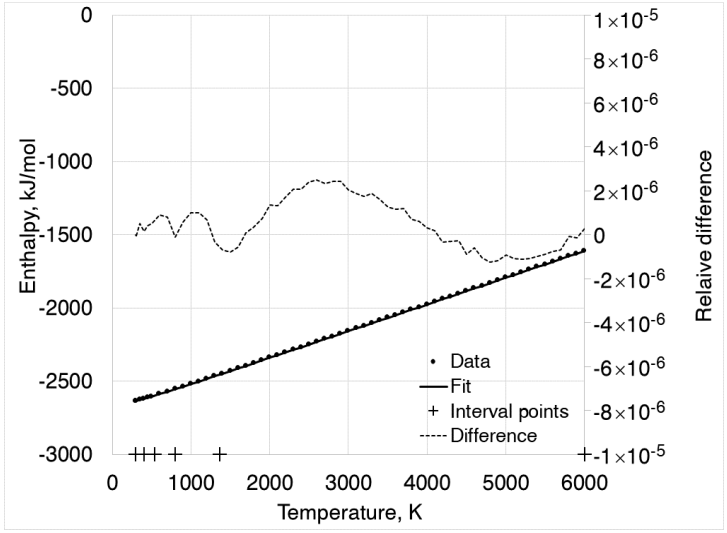


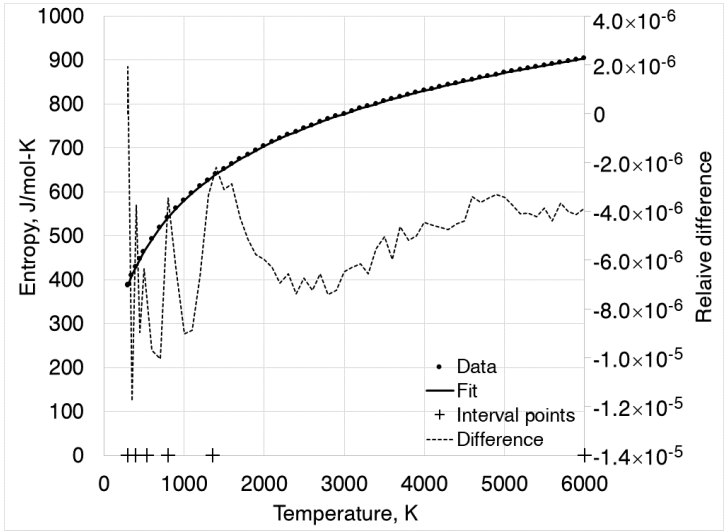


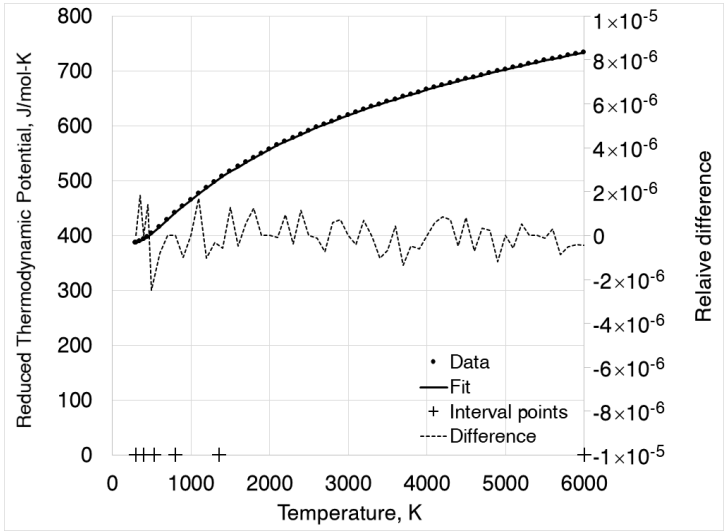
Aluminum Fluoride ($(AlF_3)_2$) – Al2F6(g)

MW = 167.9535 g/mol – $\Delta_f H^o$(298.15 K) = -2633.619 kJ/mol – S^o(298.15 K) = 387.294 J/mol-K – p^o = 0.1 MPa



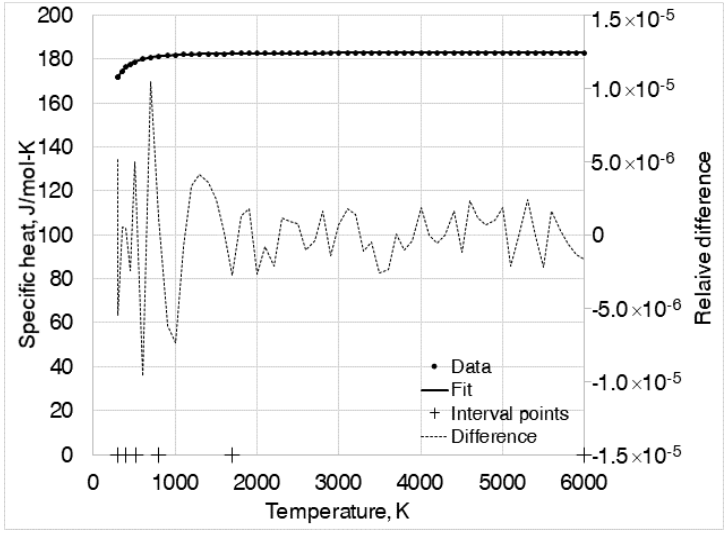


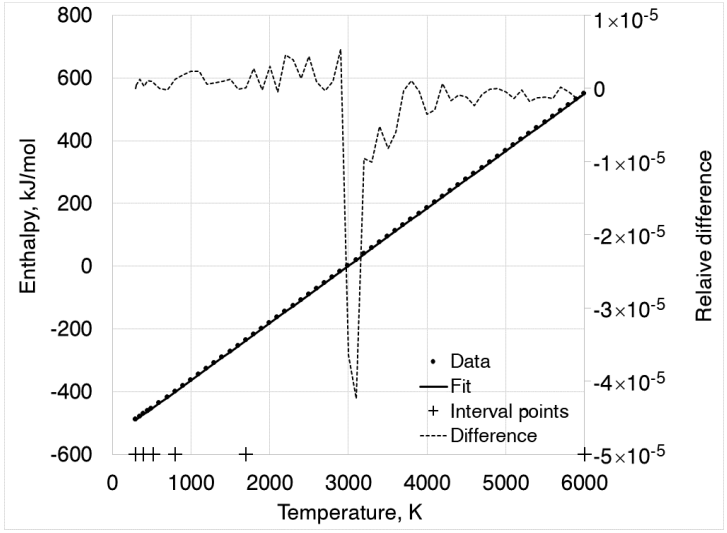


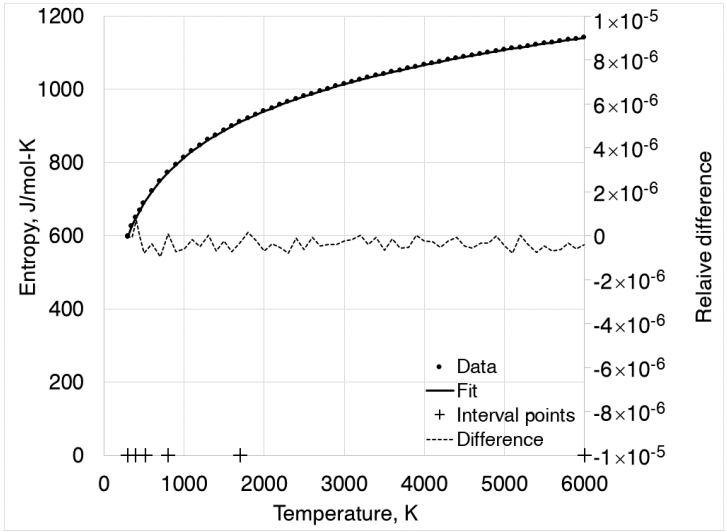


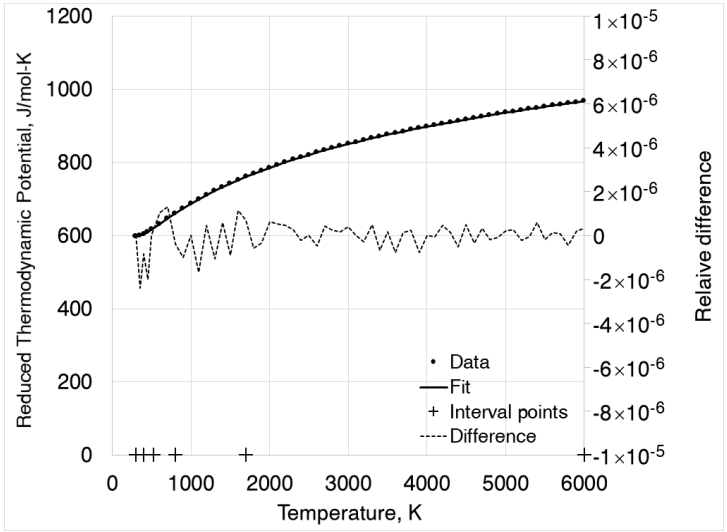
Aluminum Iodide ($(AlI_3)_2$) – Al2I6(g)

MW = 815.3901 g/mol – $\Delta_f H^o$(298.15 K) = -489.528 kJ/mol – S^o(298.15 K) = 597.38 J/mol-K – p^o = 0.1 MPa



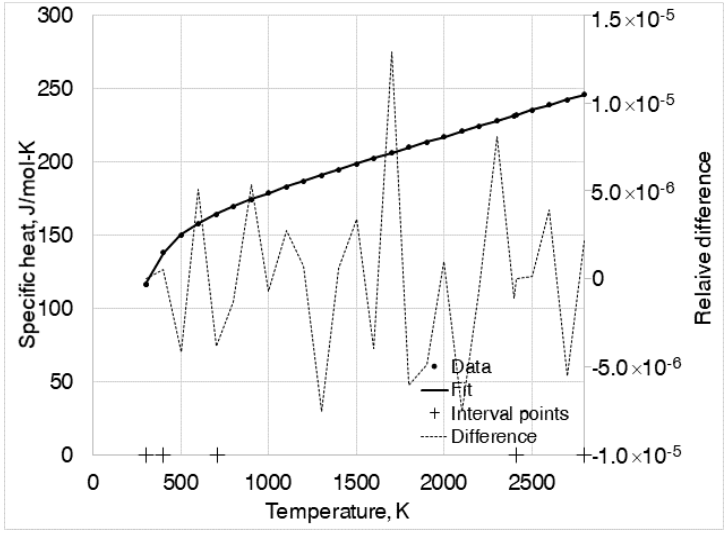


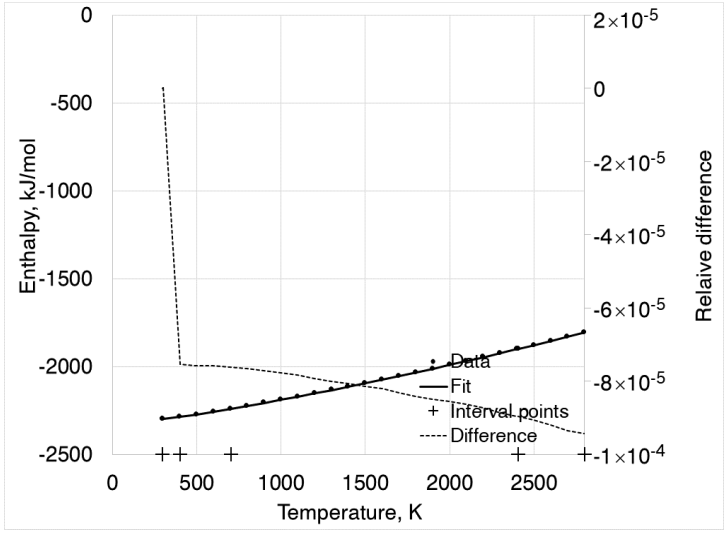


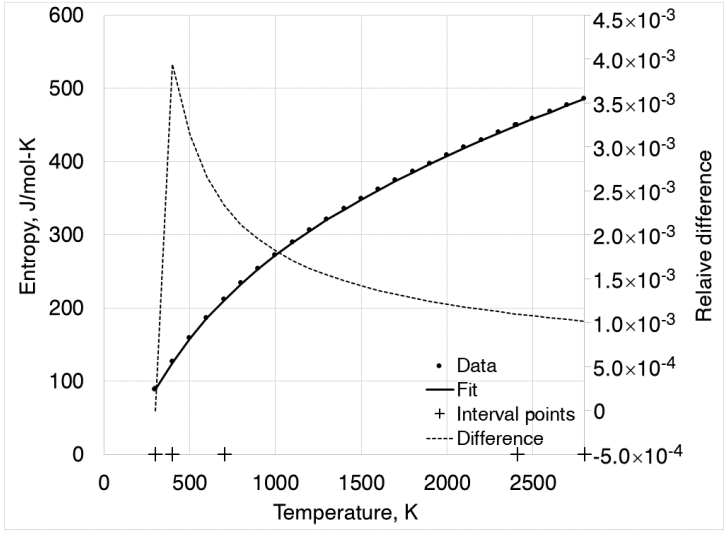


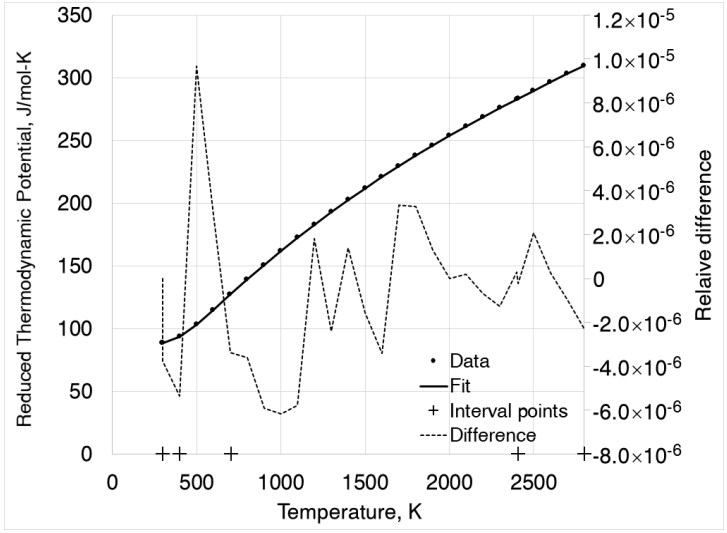
Magnesium Aluminum Oxide ($MgAl_2O_4$) – Al2Mg1O4(cr)

MW = 142.266 g/mol – $\Delta_f H^o$(298.15 K) = -2299.108 kJ/mol – S^o(298.15 K) = 88.692 J/mol-K – p^o = 0.1 MPa



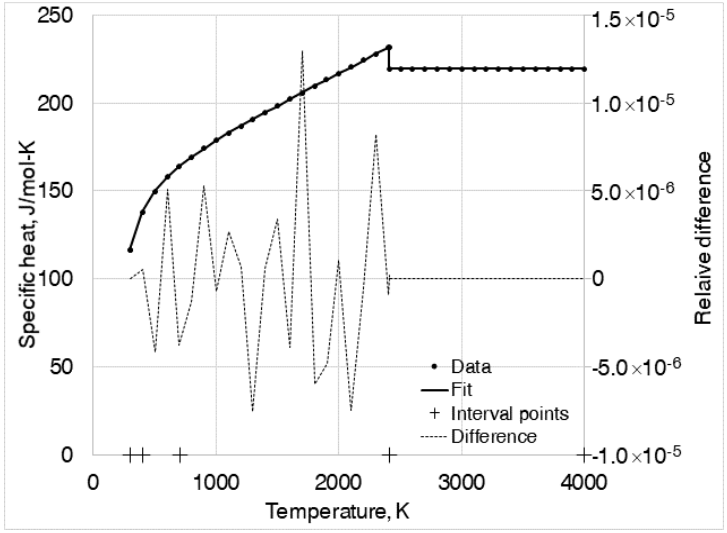


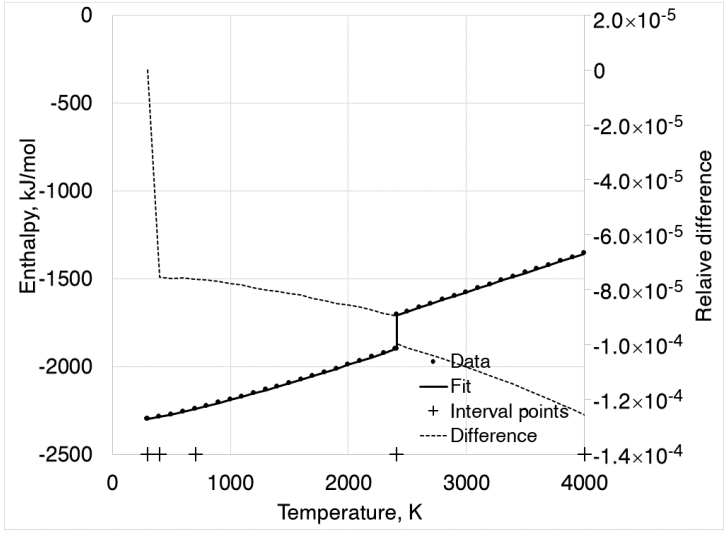


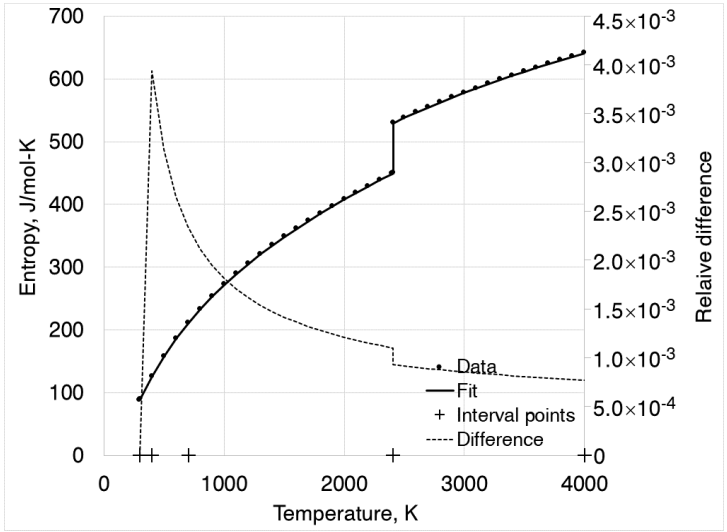


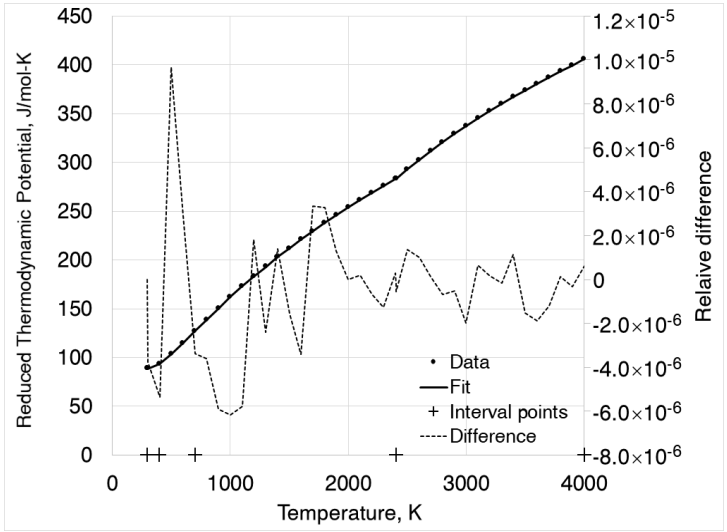
Magnesium Aluminum Oxide ($MgAl_2O_4$) – Al2Mg1O4(cr,l)

MW = 142.266 g/mol – $\Delta_f H^o$(298.15 K) = -2299.108 kJ/mol – S^o(298.15 K) = 88.692 J/mol-K – p^o = 0.1 MPa



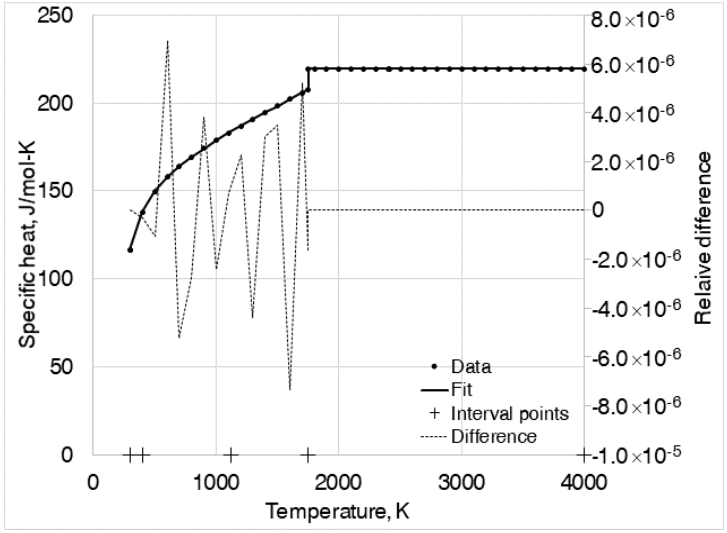


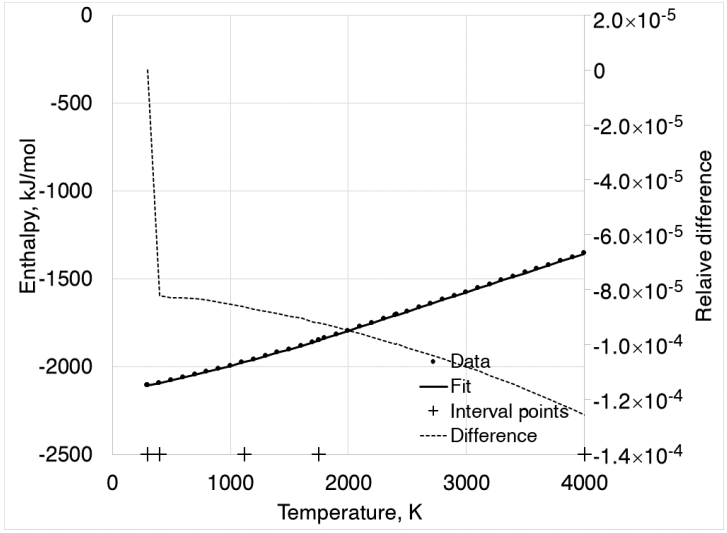


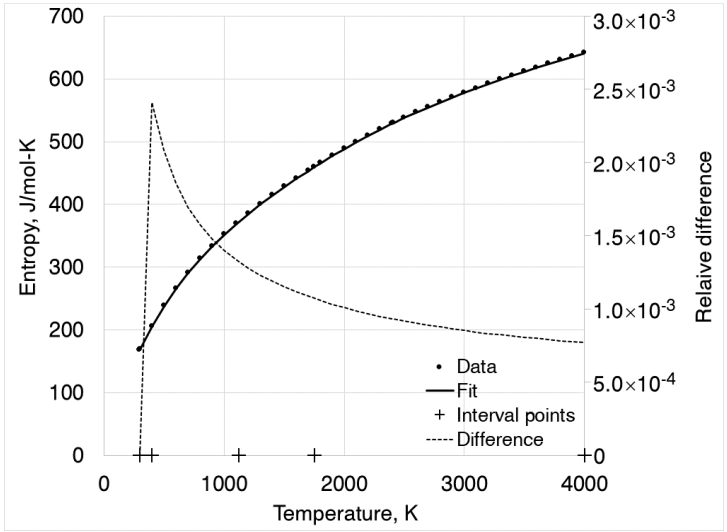


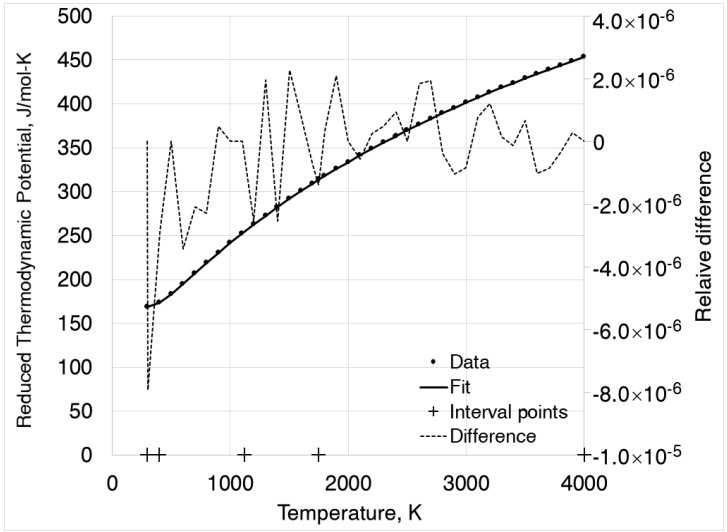
Magnesium Aluminum Oxide ($MgAl_2O_4$) – Al2Mg1O4(l)

MW = 142.266 g/mol – $\Delta_f H^o$(298.15 K) = -2106.527 kJ/mol – S^o(298.15 K) = 168.473 J/mol-K – p^o = 0.1 MPa



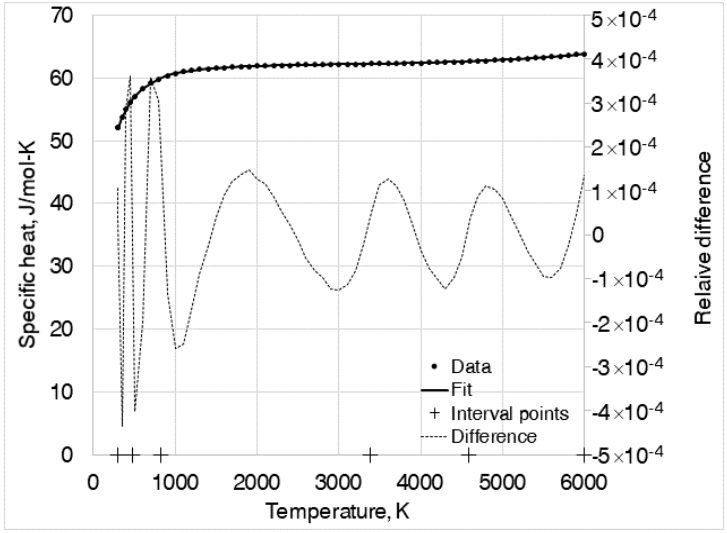


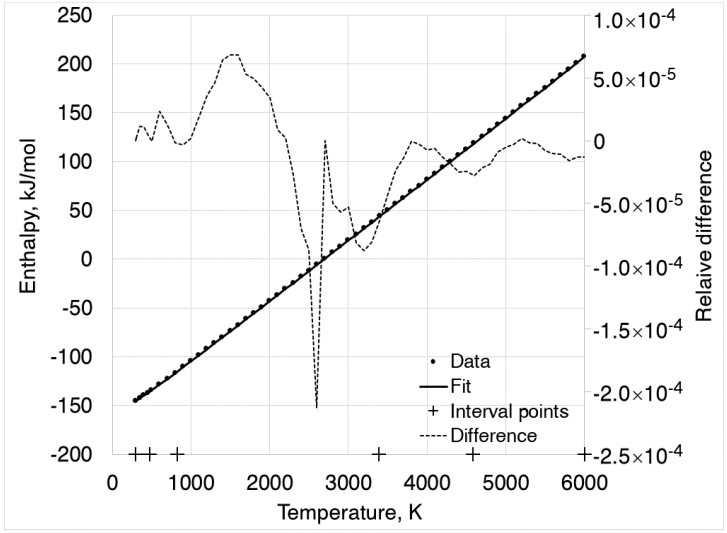


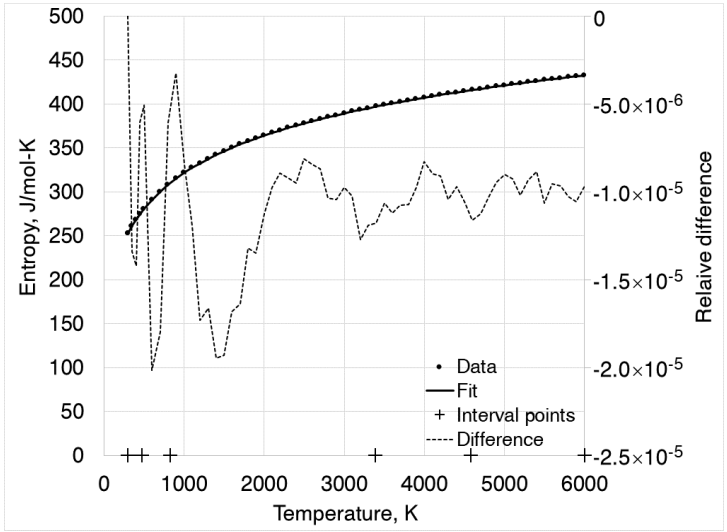


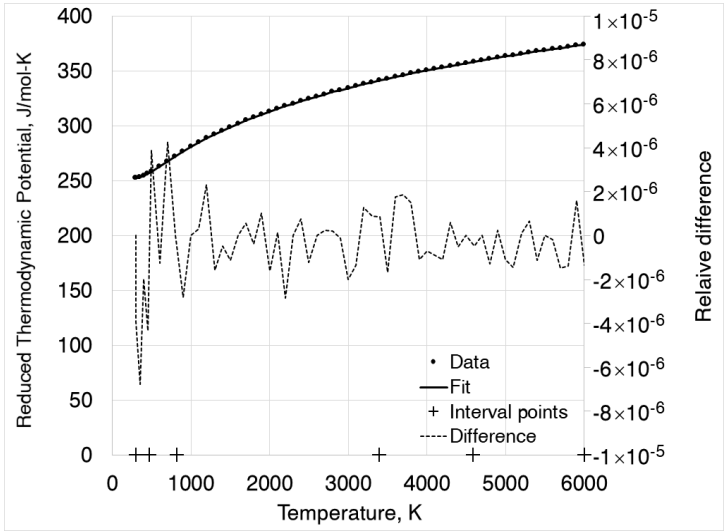
Aluminum Oxide (Al_2O) – Al2O1(g)

MW = 69.9625 g/mol – $\Delta_f H^o$(298.15 K) = -145.185 kJ/mol – S^o(298.15 K) = 252.332 J/mol-K – p^o = 0.1 MPa



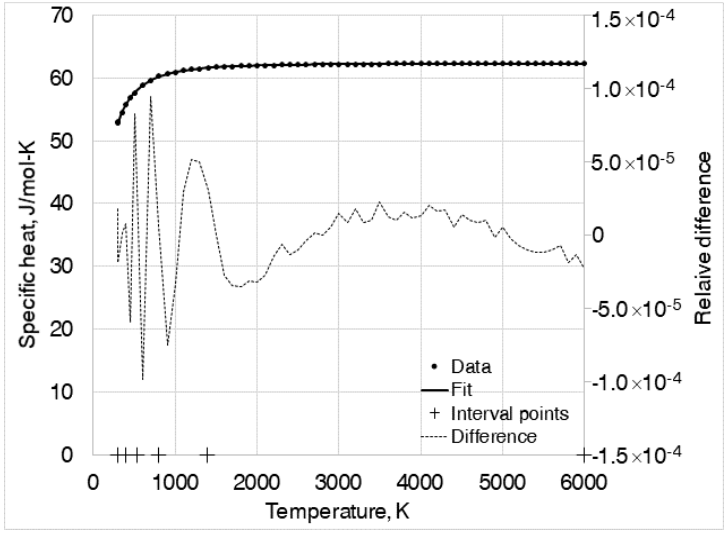


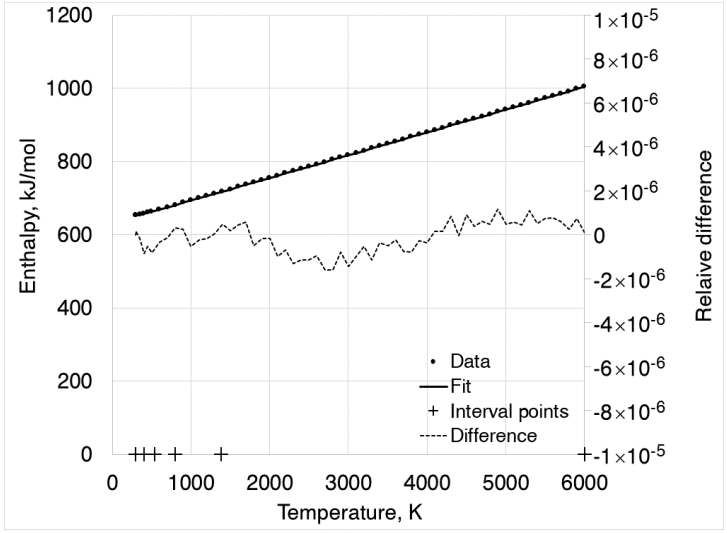


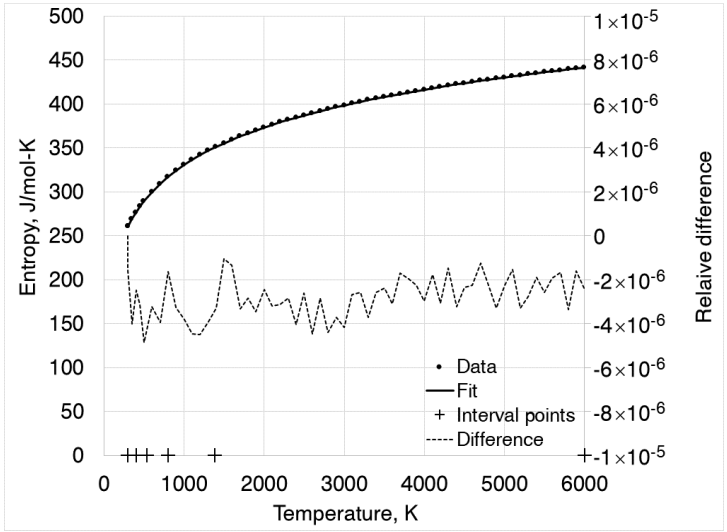


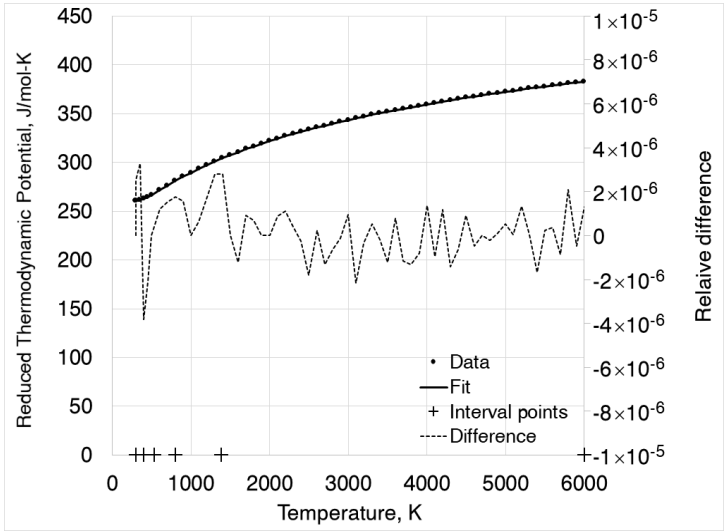
Aluminum Oxide, Ion (Al_2O^+) – Al2O1+(g)

MW = 69.9619 g/mol – $\Delta_f H^o$(298.15 K) = 652.436 kJ/mol – S^o(298.15 K) = 260.66 J/mol-K – p^o = 0.1 MPa



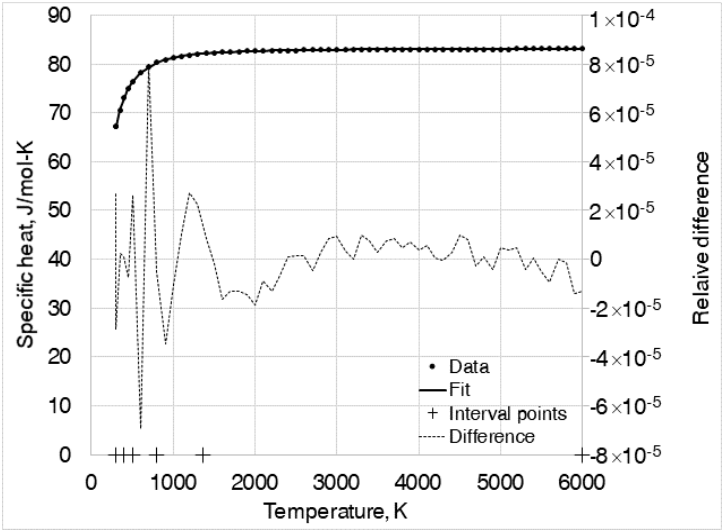


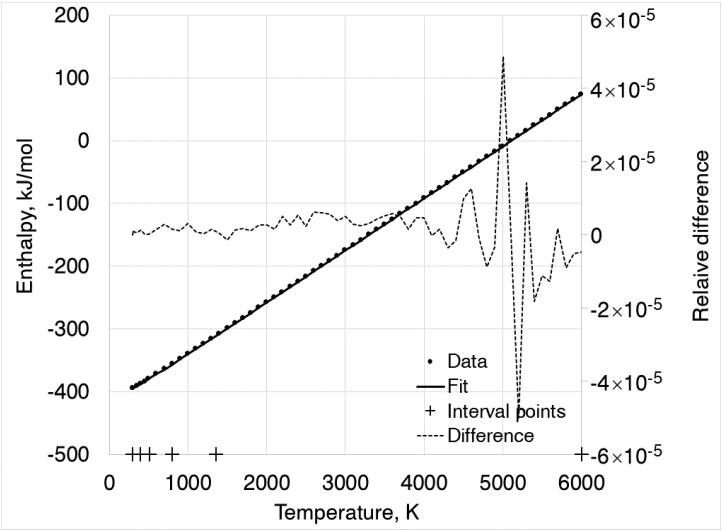


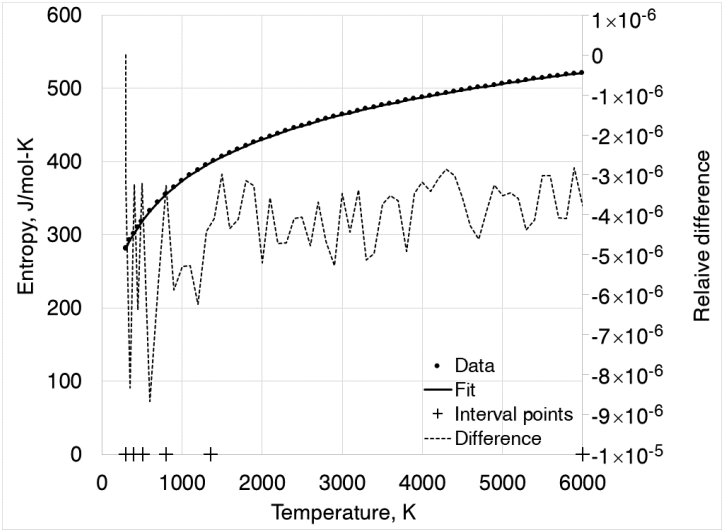


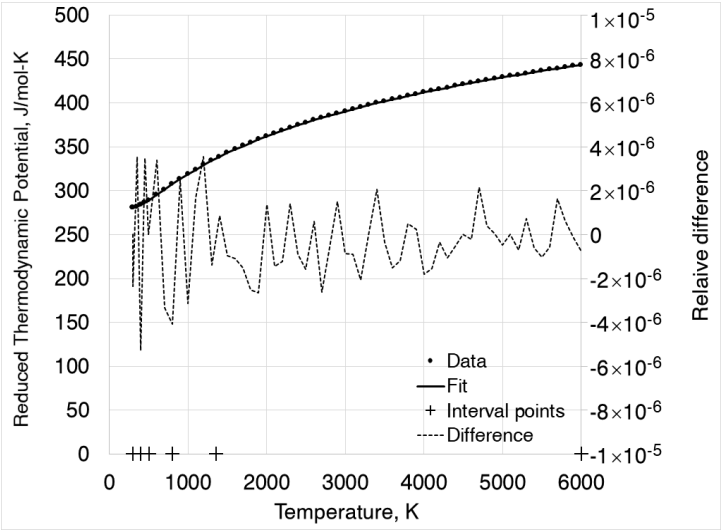
Aluminum Oxide ($(AlO)_2$) – Al2O2(g)

MW = 85.9619 g/mol – $\Delta_f H^o$(298.15 K) = -394.551 kJ/mol – S^o(298.15 K) = 280.98 J/mol-K – p^o = 0.1 MPa



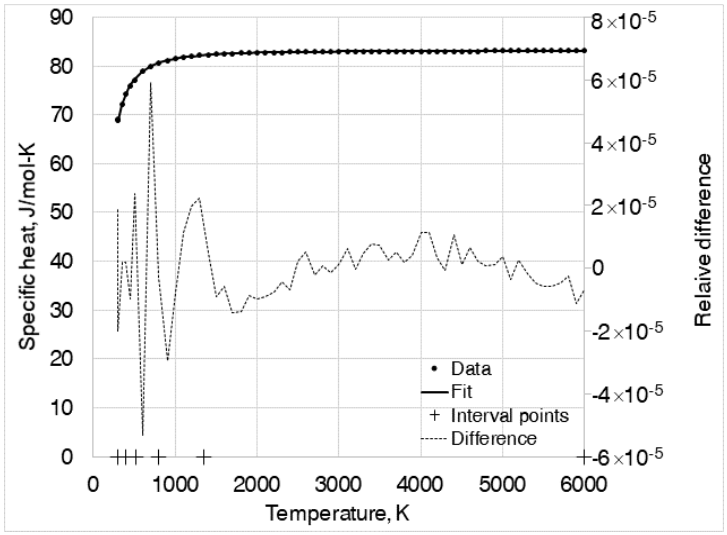


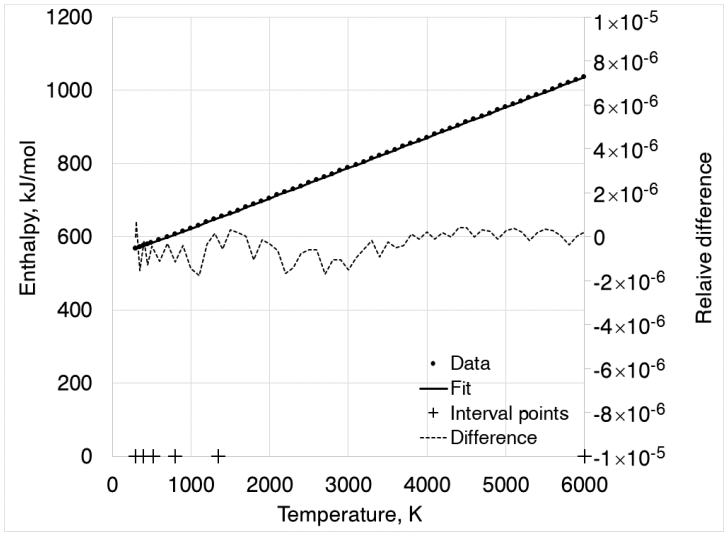


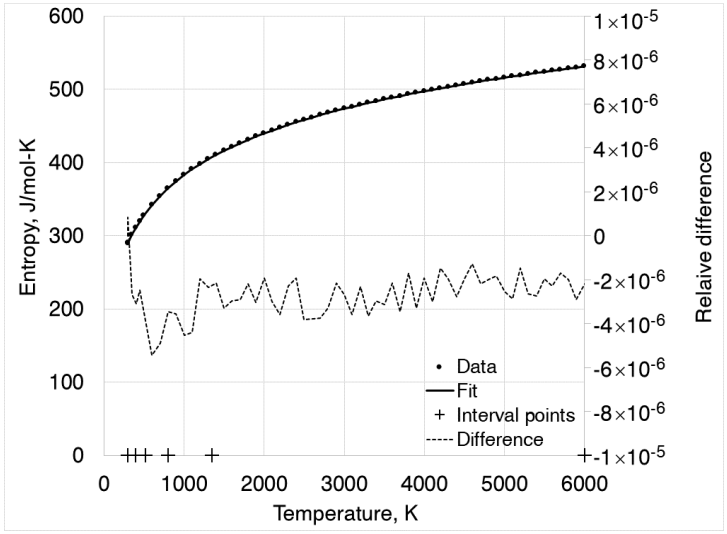


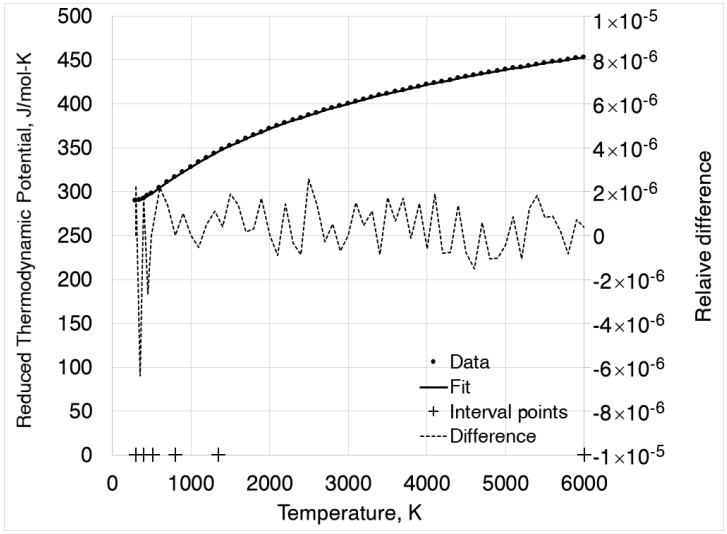
Aluminum Oxide, Ion ($Al_2O_2^+$) – Al2O2+(g)

MW = 85.9613 g/mol – $\Delta_f H^o$(298.15 K) = 567.414 kJ/mol – S^o(298.15 K) = 289.7 J/mol-K – p^o = 0.1 MPa



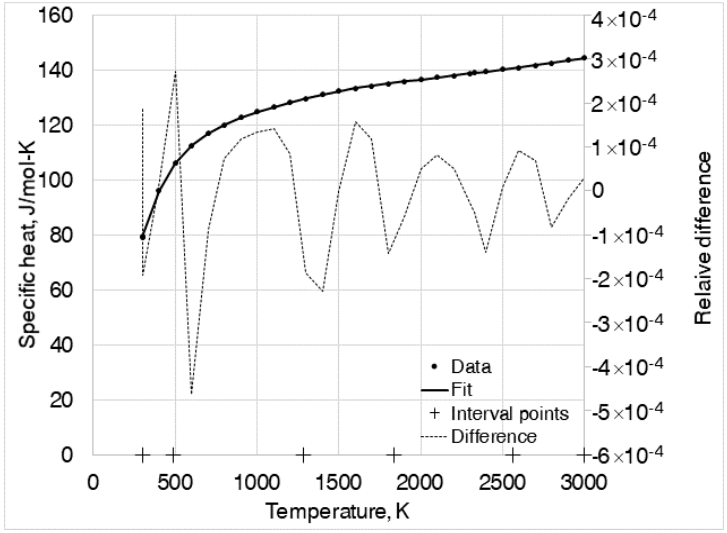


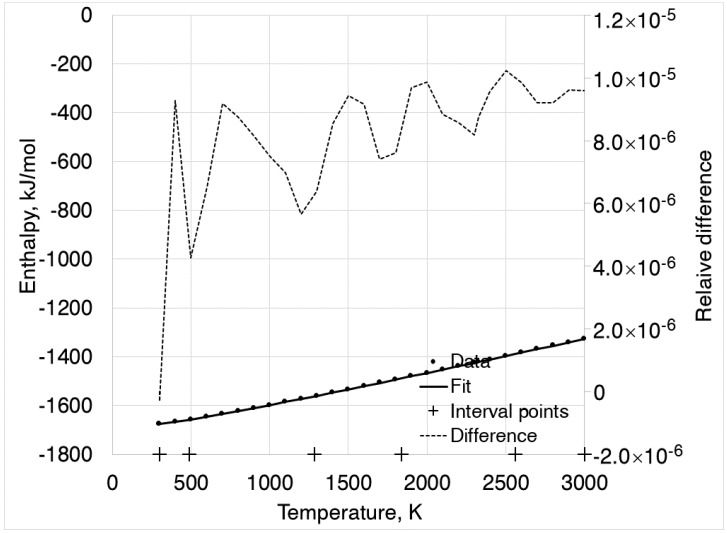


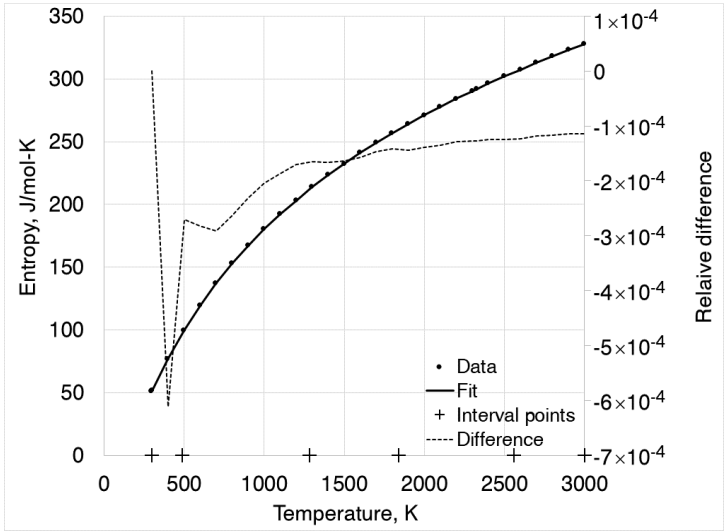


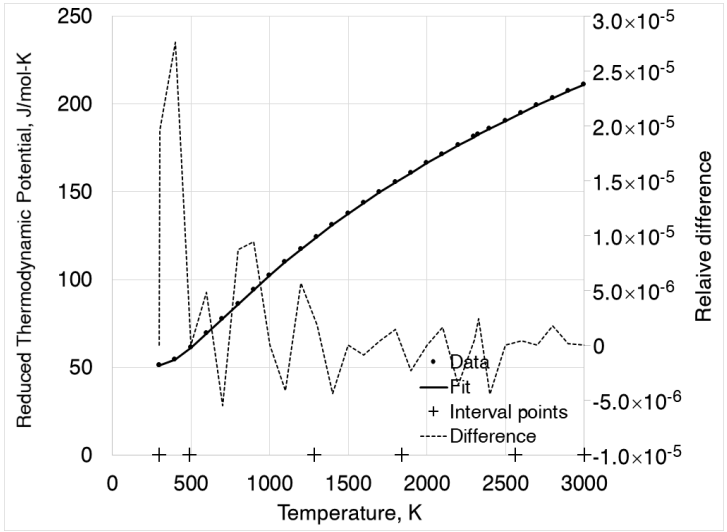
Aluminum Oxide, Alpha (Al_2O_3) – Al2O3(cr)_01

MW = 101.9613 g/mol – $\Delta_f H^o$(298.15 K) = -1675.692 kJ/mol – S^o(298.15 K) = 50.95 J/mol-K – p^o = 0.1 MPa



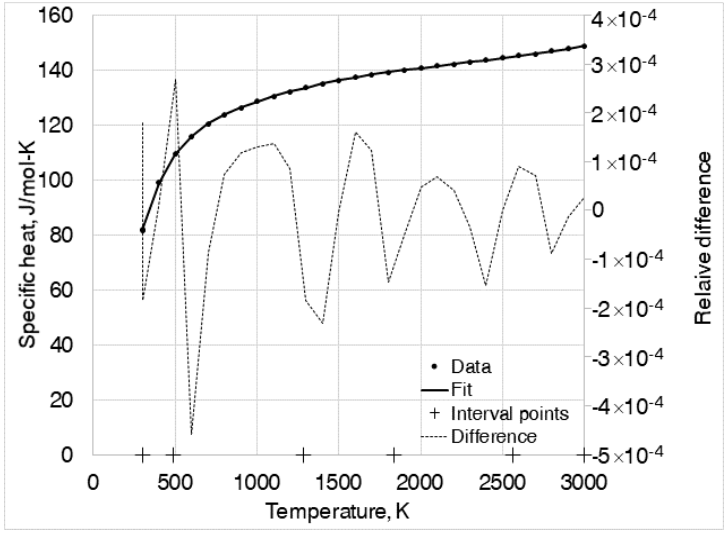


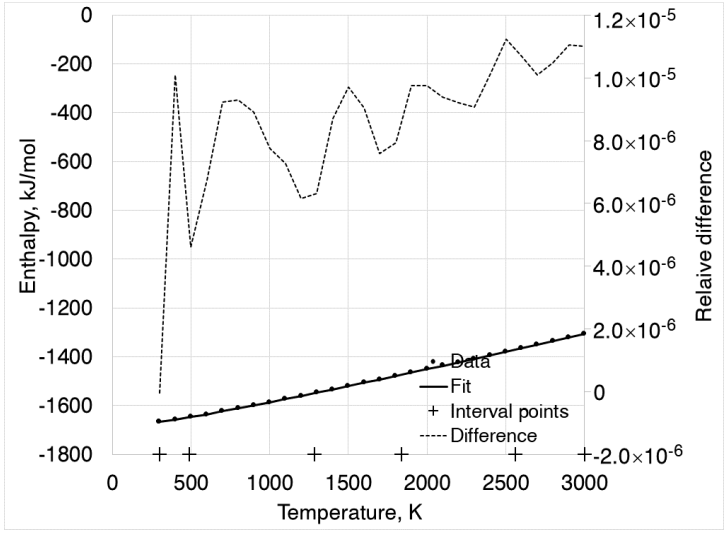


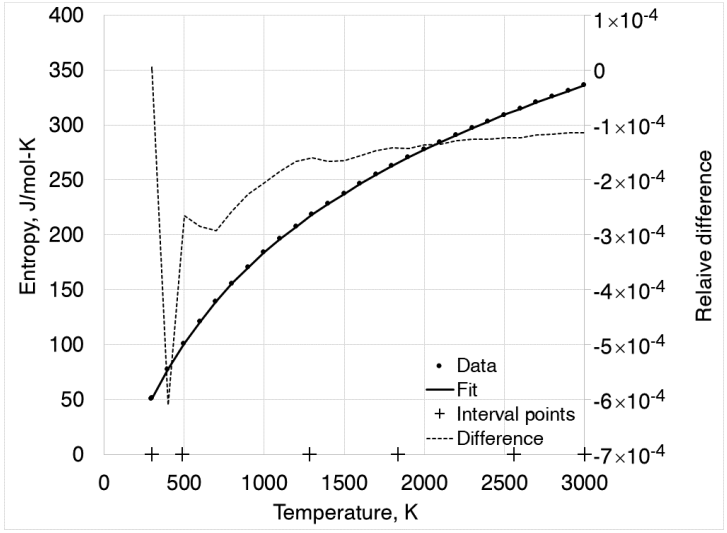


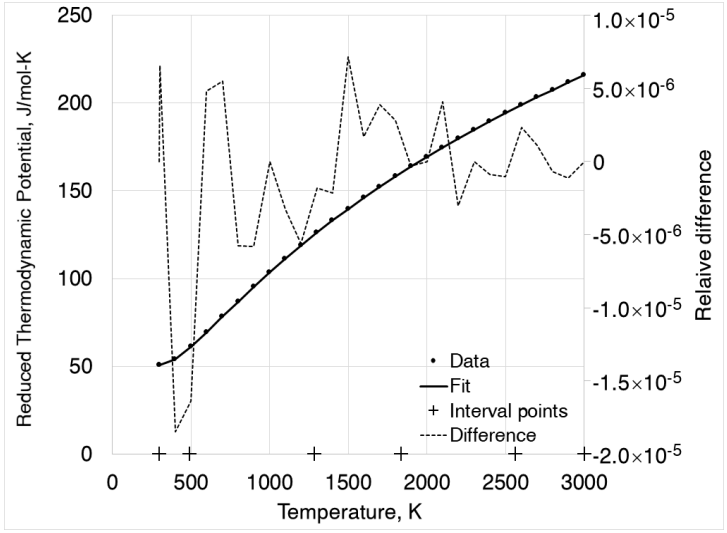
Aluminum Oxide, Delta (Al_2O_3) – Al2O3(cr)_02

MW = 101.9613 g/mol – $\Delta_f H^o$(298.15 K) = -1666.487 kJ/mol – S^o(298.15 K) = 50.626 J/mol-K – p^o = 0.1 MPa



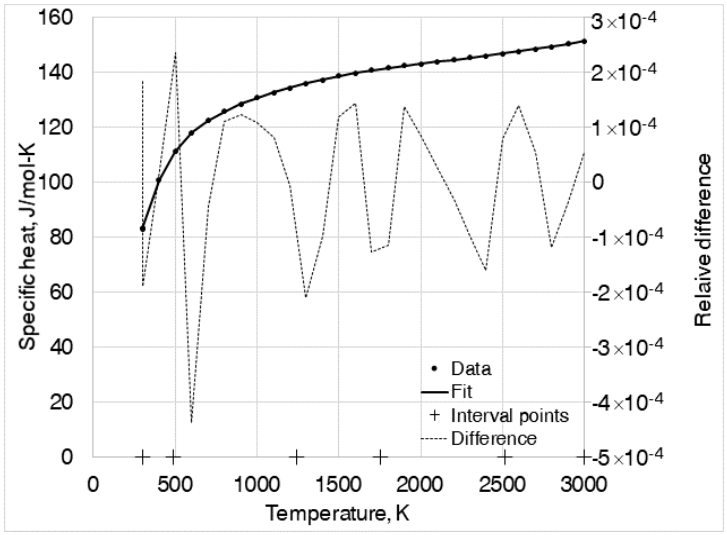


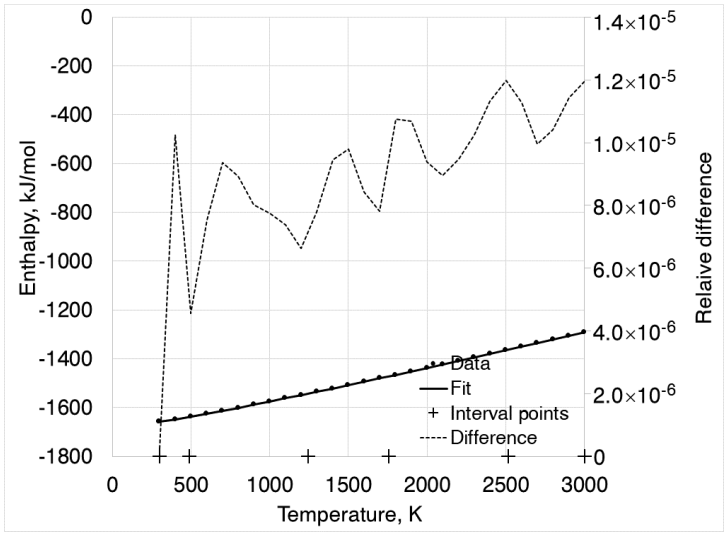


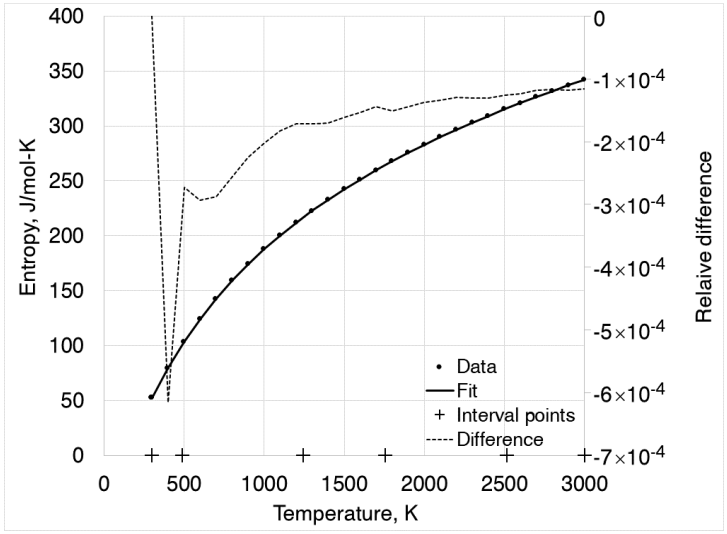


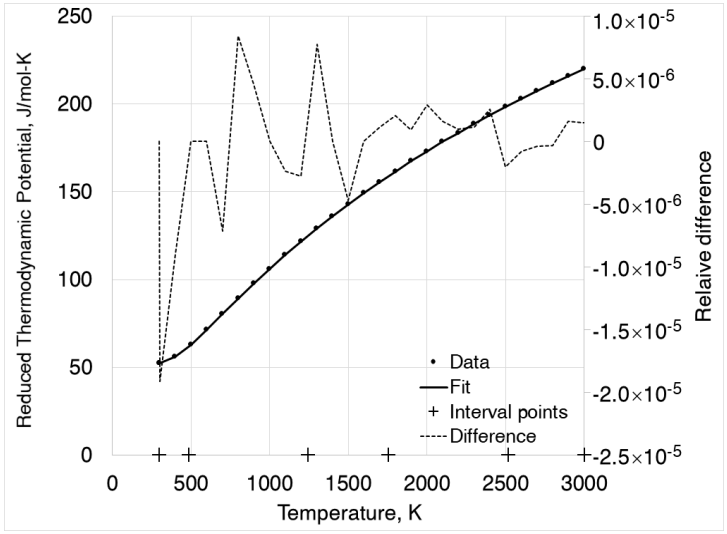
Aluminum Oxide, Gamma (Al_2O_3) – Al2O3(cr)_03

MW = 101.9613 g/mol – $\Delta_f H^o$(298.15 K) = -1656.864 kJ/mol – S^o(298.15 K) = 52.3 J/mol-K – p^o = 0.1 MPa



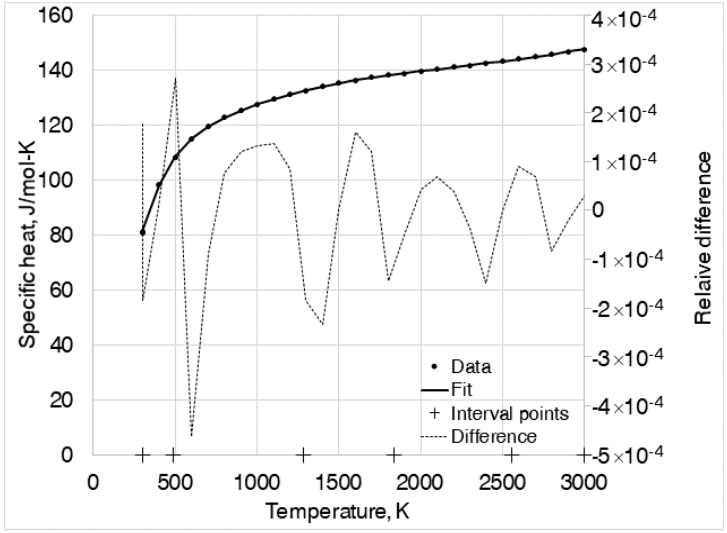


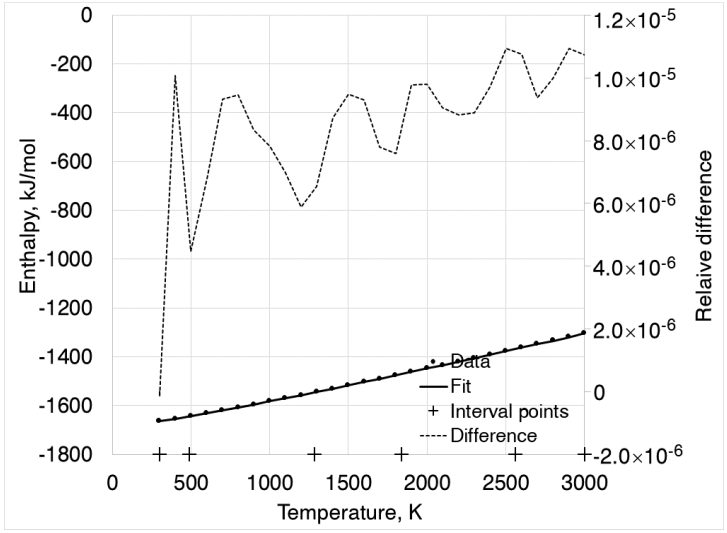


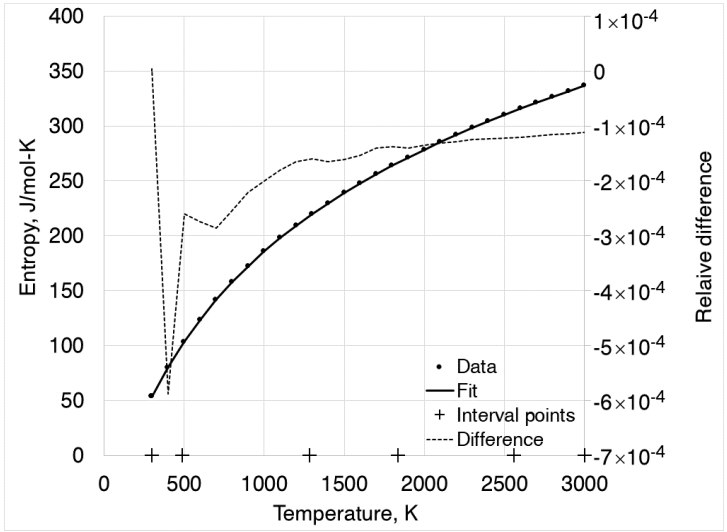


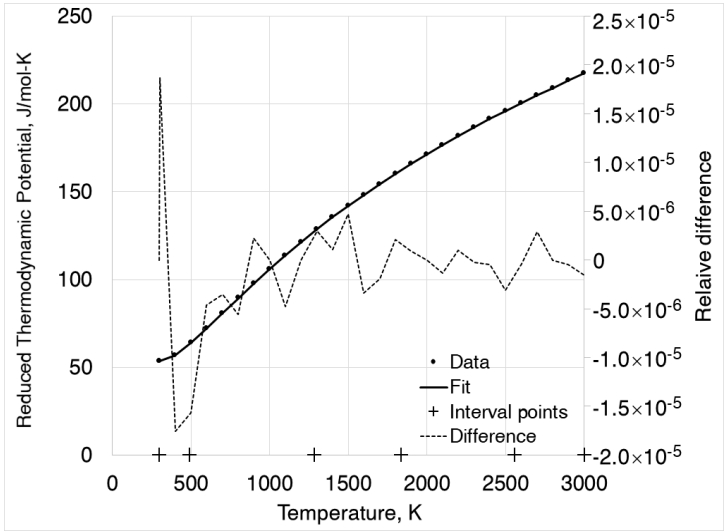
Aluminum Oxide, Kappa (Al_2O_3) – Al2O3(cr)_04

MW = 101.9613 g/mol – $\Delta_f H^o$(298.15 K) = -1662.303 kJ/mol – S^o(298.15 K) = 53.555 J/mol-K – p^o = 0.1 MPa



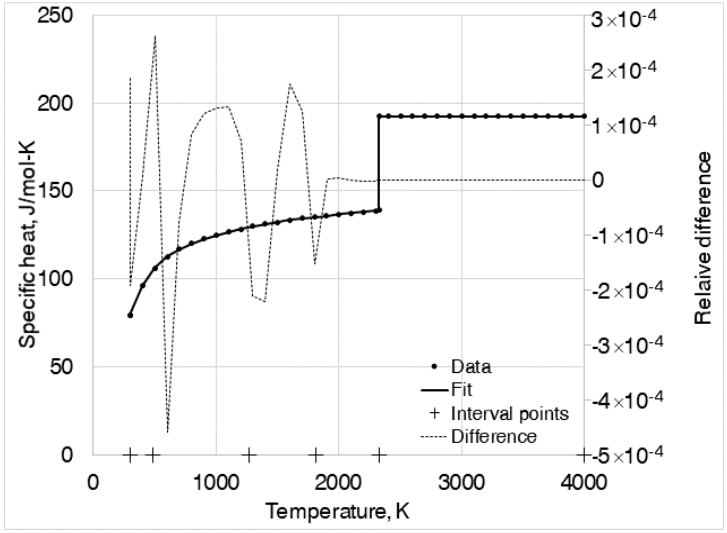


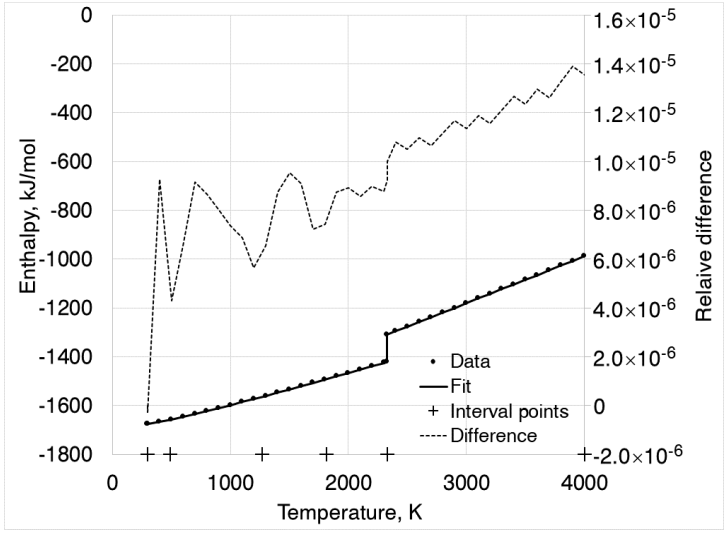


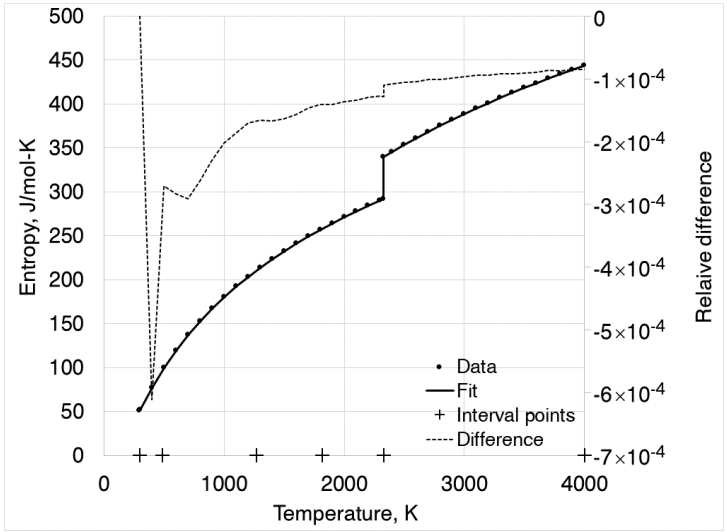


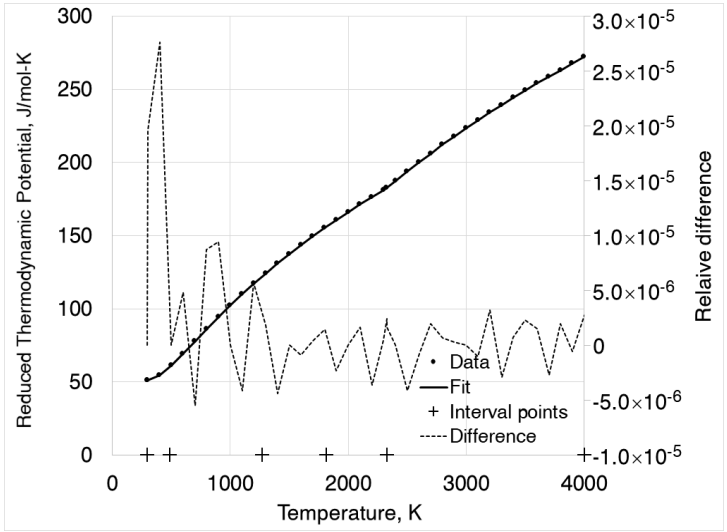
Aluminum Oxide (Al_2O_3) – Al2O3(cr,l)

MW = 101.9613 g/mol – $\Delta_f H^o$(298.15 K) = -1675.692 kJ/mol – S^o(298.15 K) = 50.95 J/mol-K – p^o = 0.1 MPa



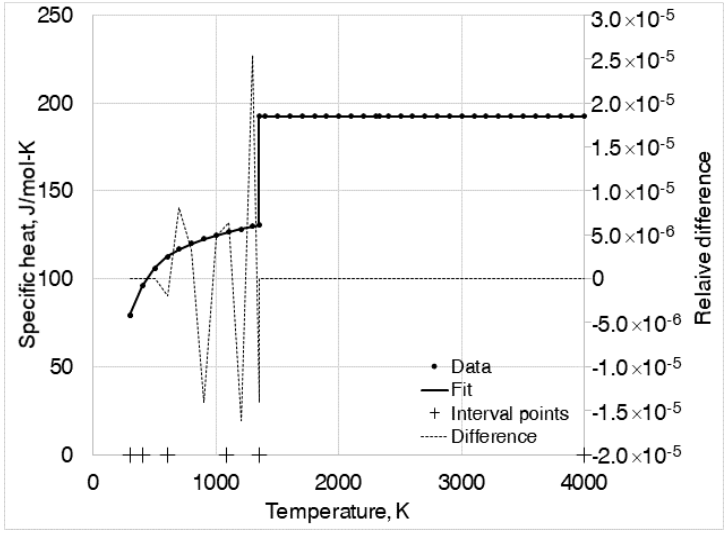


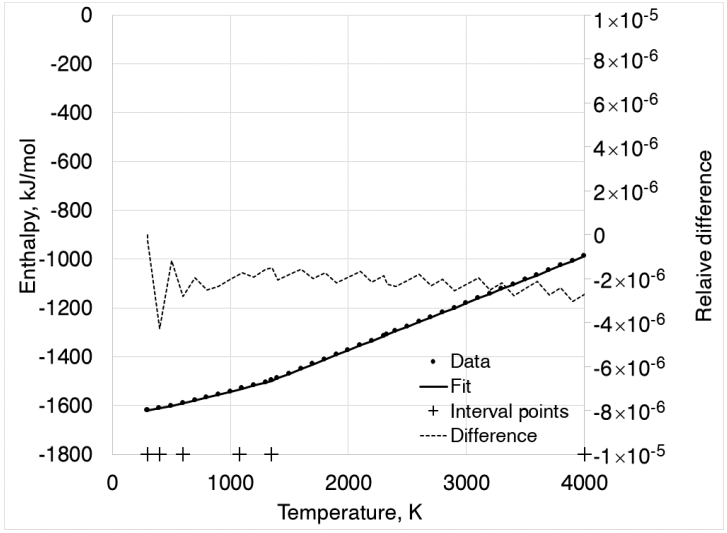


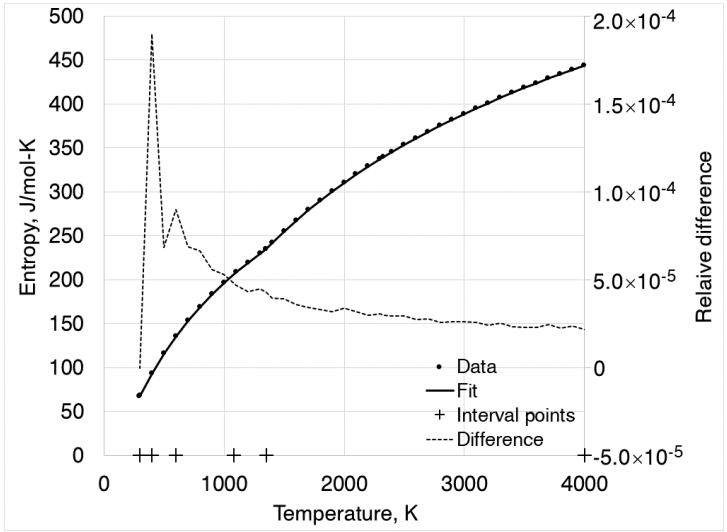


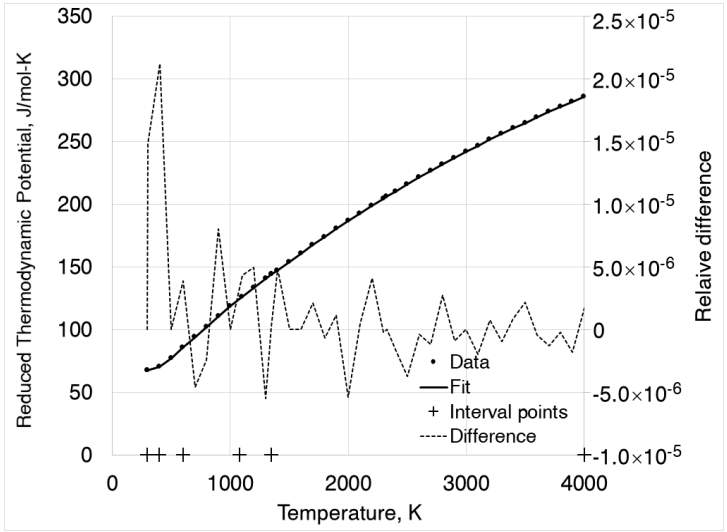
Aluminum Oxide (Al_2O_3) – Al2O3(l)

MW = 101.9613 g/mol – $\Delta_f H^o$(298.15 K) = -1620.567 kJ/mol – S^o(298.15 K) = 67.298 J/mol-K – p^o = 0.1 MPa



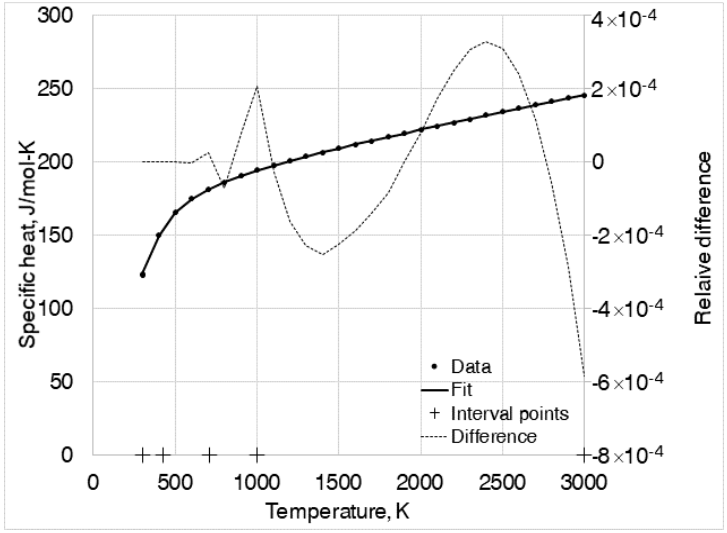


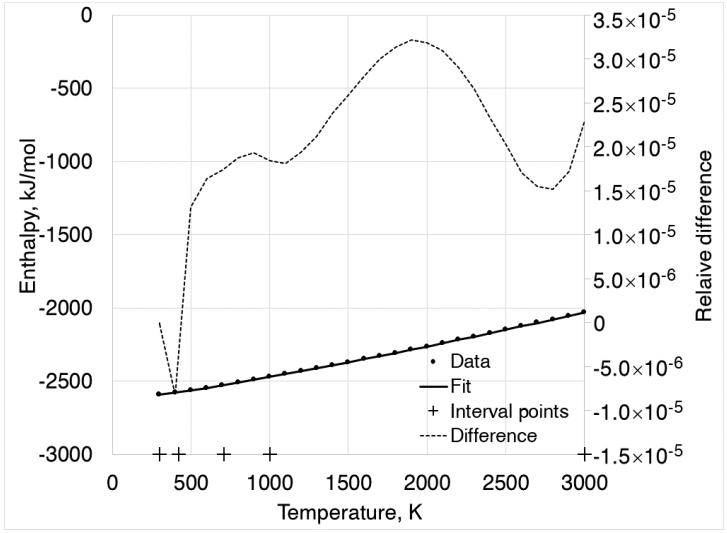


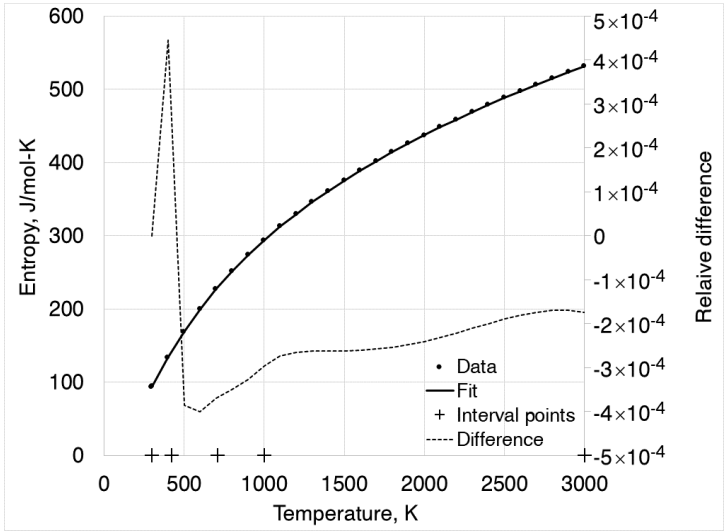


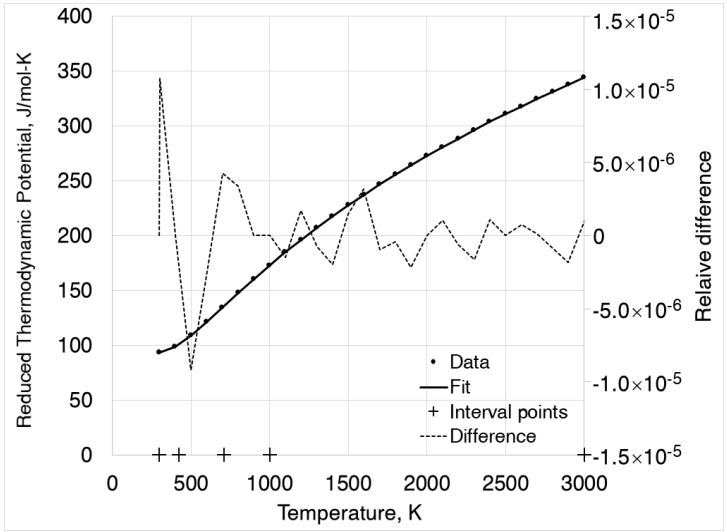
Aluminum Silicate, Andalusite (Al_2SiO_5) – Al2O5Si1(cr)_01

MW = 162.0456 g/mol – $\Delta_f H^o$(298.15 K) = -2592.072 kJ/mol – S^o(298.15 K) = 93.221 J/mol-K – p^o = 0.1 MPa



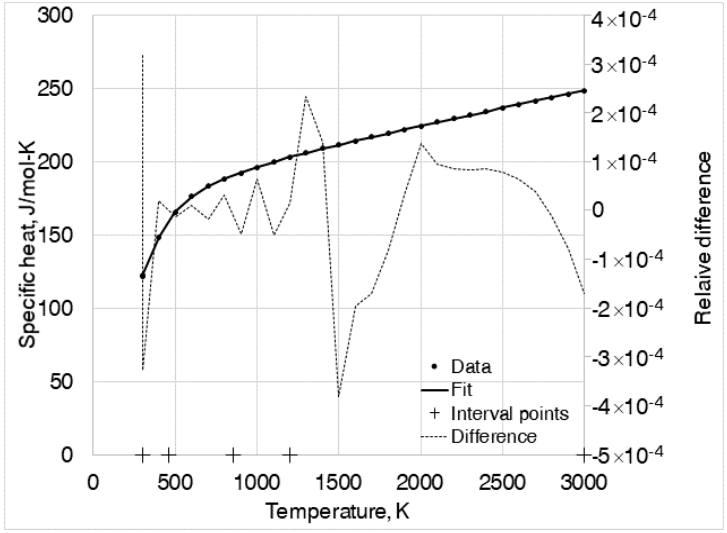


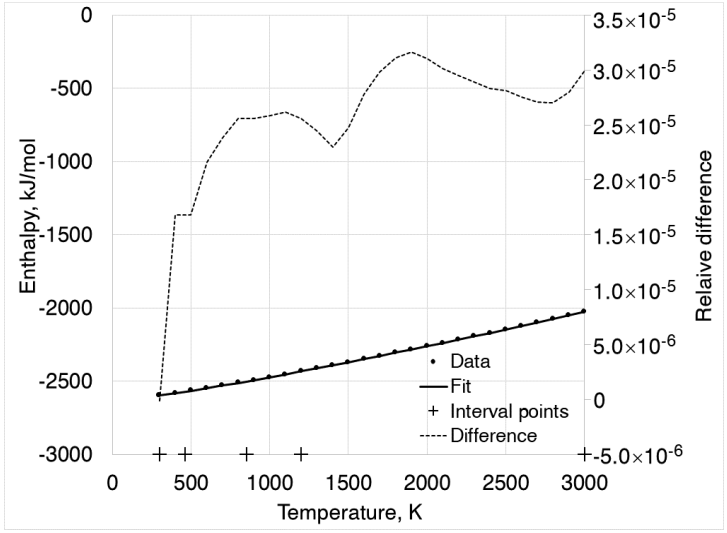


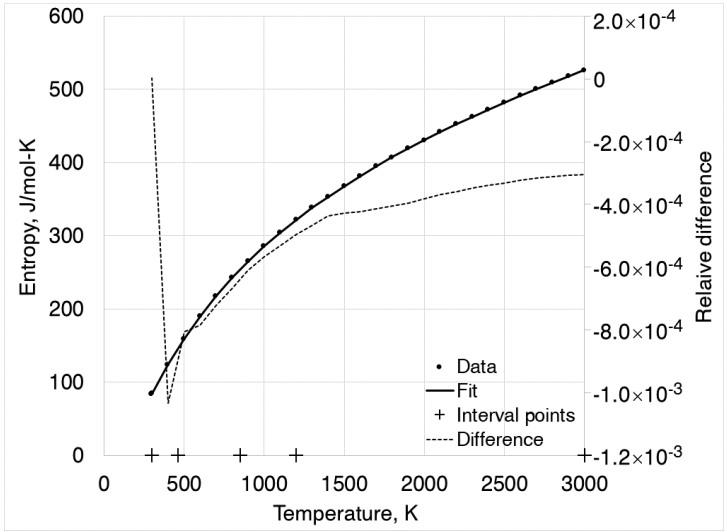


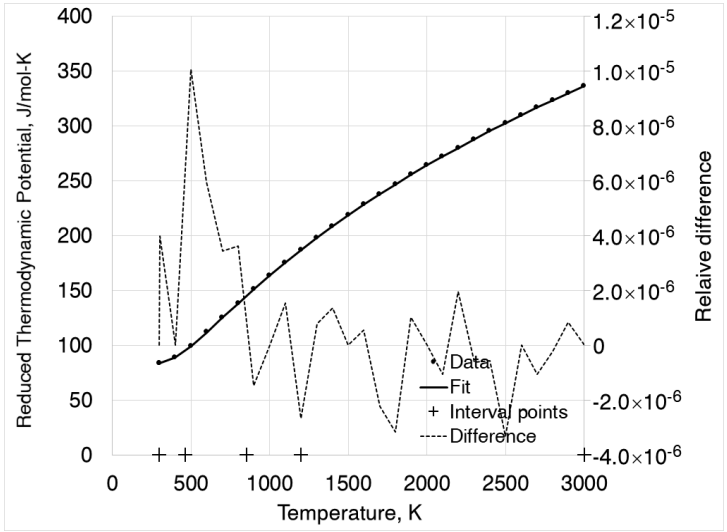
Aluminum Silicate, Kyanite (Al_2SiO_5) – Al2O5Si1(cr)_02

MW = 162.0456 g/mol – $\Delta_f H^o$(298.15 K) = -2594.331 kJ/mol – S^o(298.15 K) = 83.815 J/mol-K – p^o = 0.1 MPa



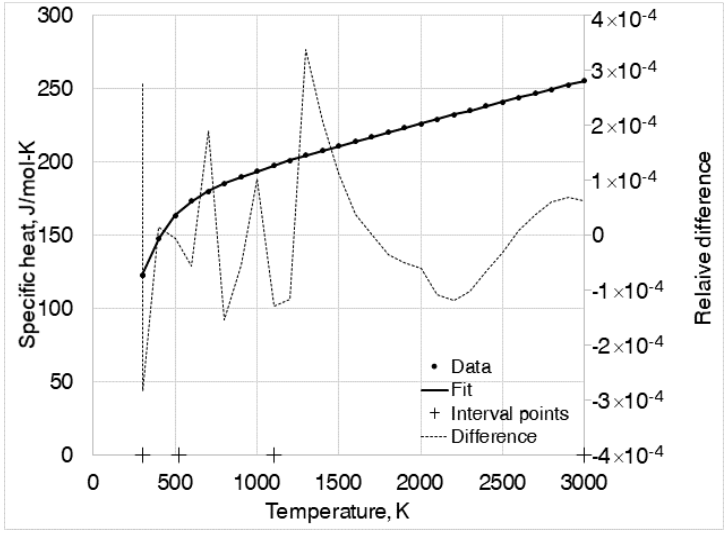


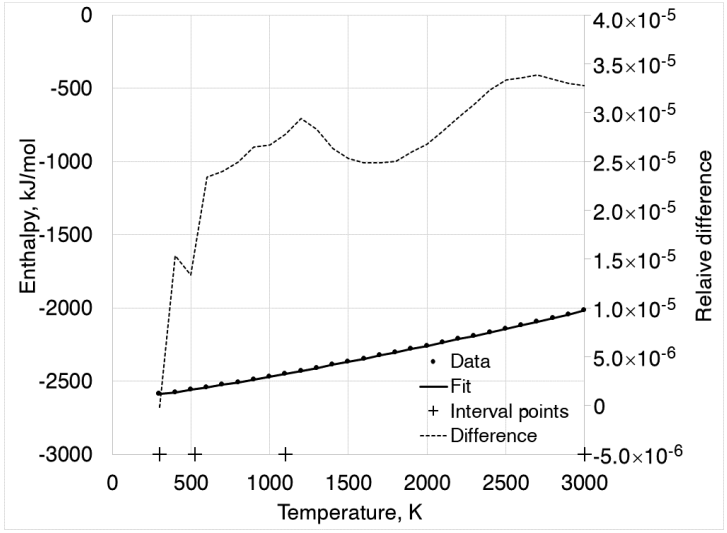


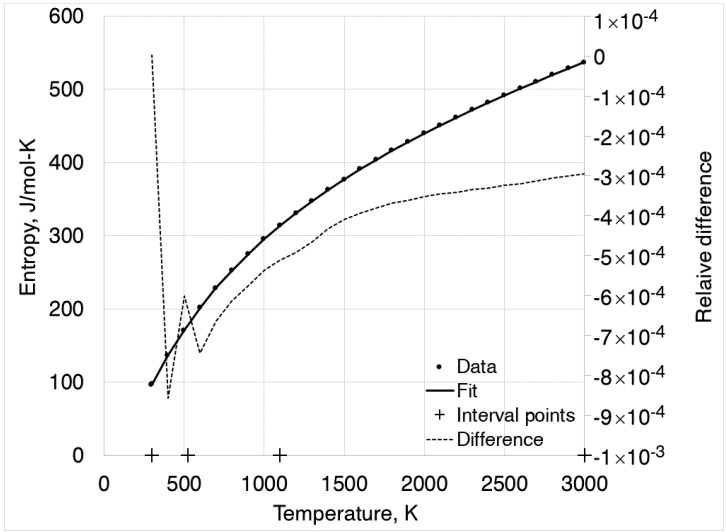


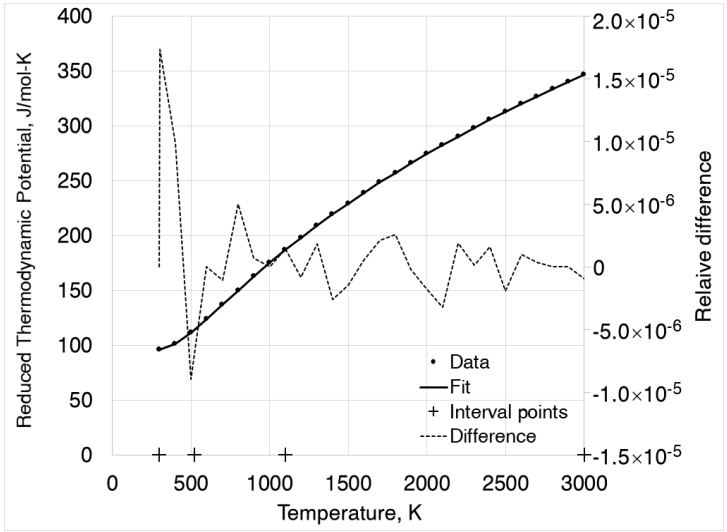
Aluminum Silicate, Sillimanite (Al_2SiO_5) – Al2O5Si1(cr)_03

MW = 162.0456 g/mol – $\Delta_f H^o$(298.15 K) = -2589.101 kJ/mol – S^o(298.15 K) = 96.179 J/mol-K – p^o = 0.1 MPa



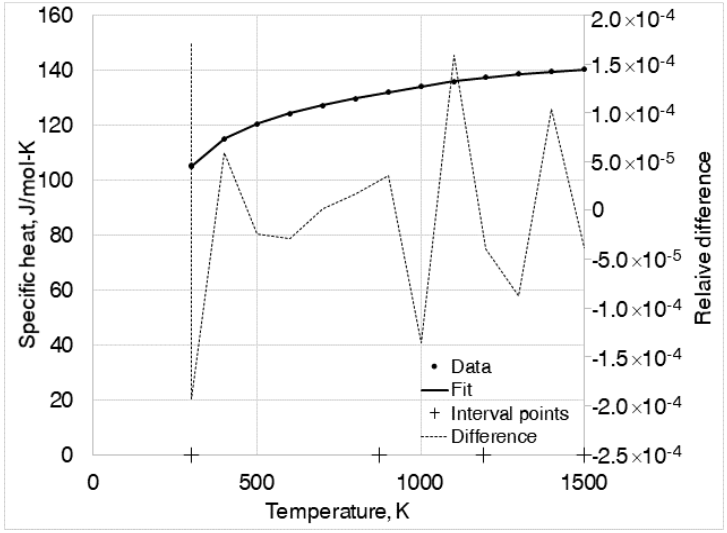


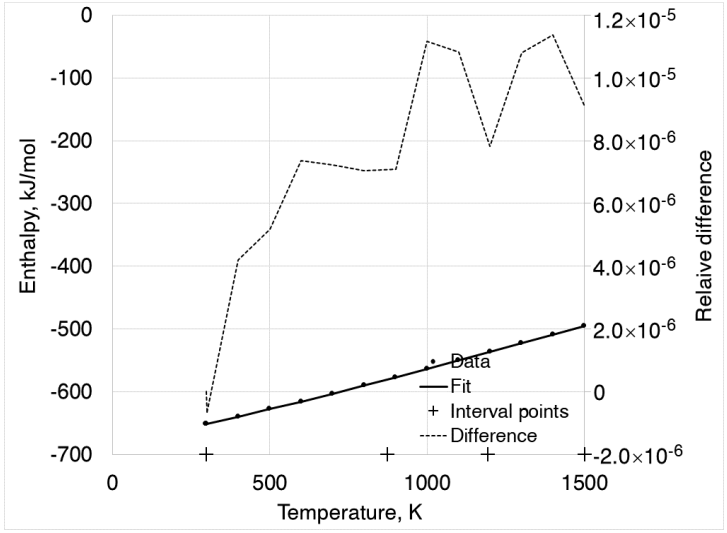


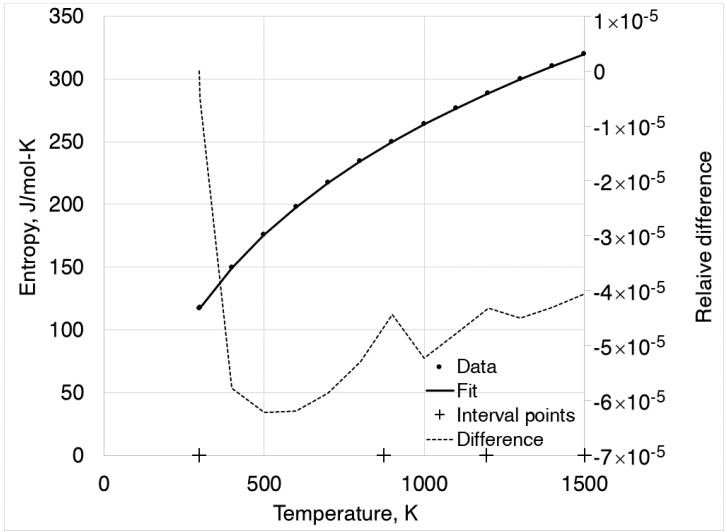


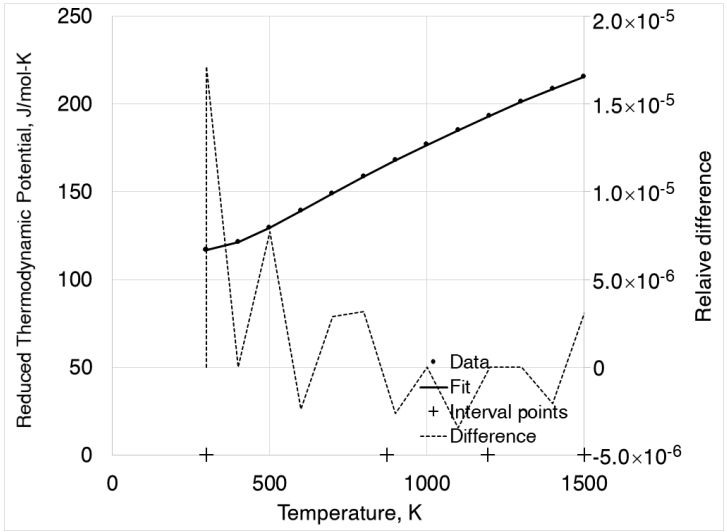
Aluminum Sulfide (Al_2S_3) – Al2S3(cr)

MW = 150.14 g/mol – $\Delta_f H^o$(298.15 K) = -651.03 kJ/mol – S^o(298.15 K) = 116.847 J/mol-K – p^o = 0.1 MPa



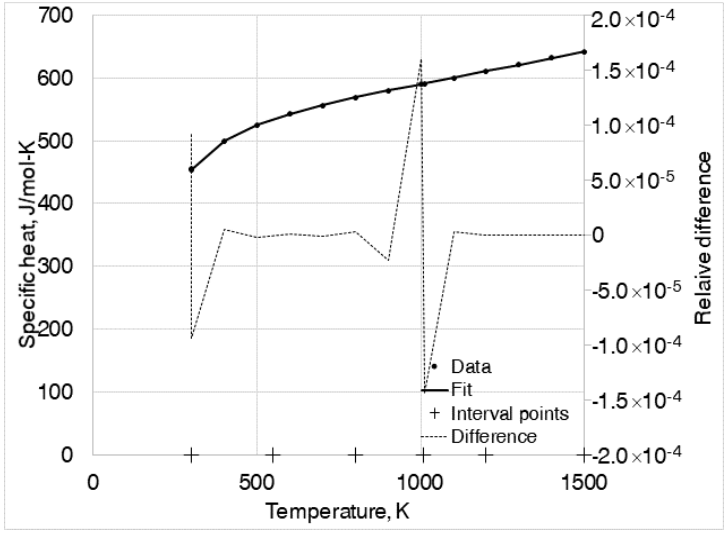


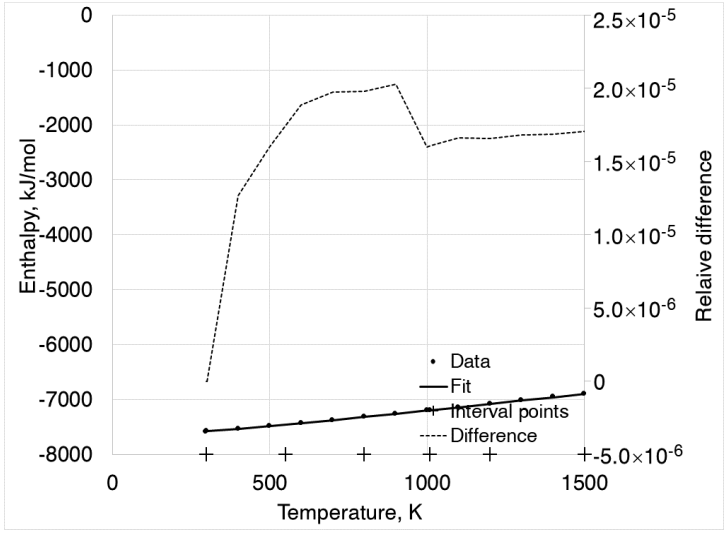


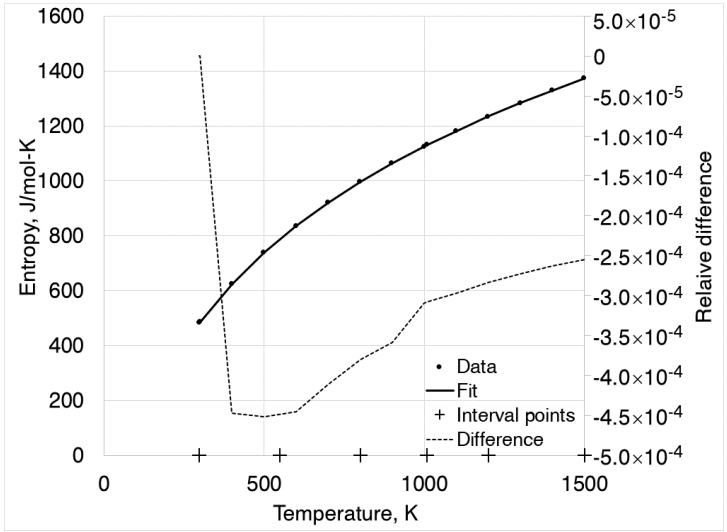


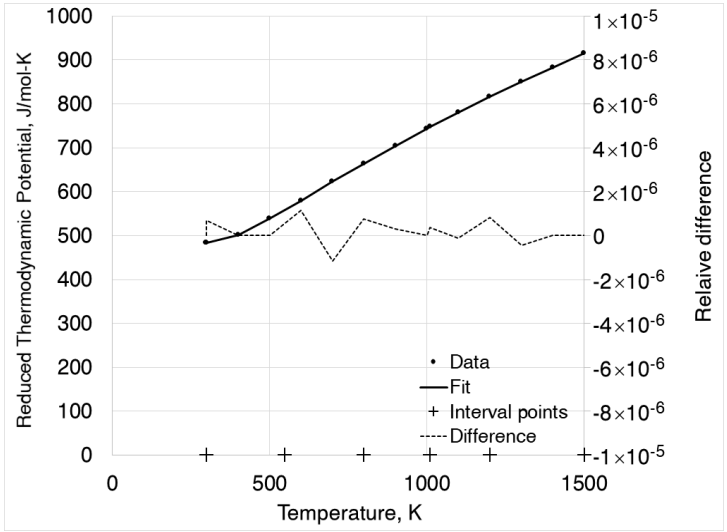
Chiolite ($Na_5Al_3F_{14}$) – Al3F14Na5(cr)

MW = 461.8711 g/mol – $\Delta_f H^o$(298.15 K) = -7581.408 kJ/mol – S^o(298.15 K) = 483.545 J/mol-K – p^o = 0.1 MPa



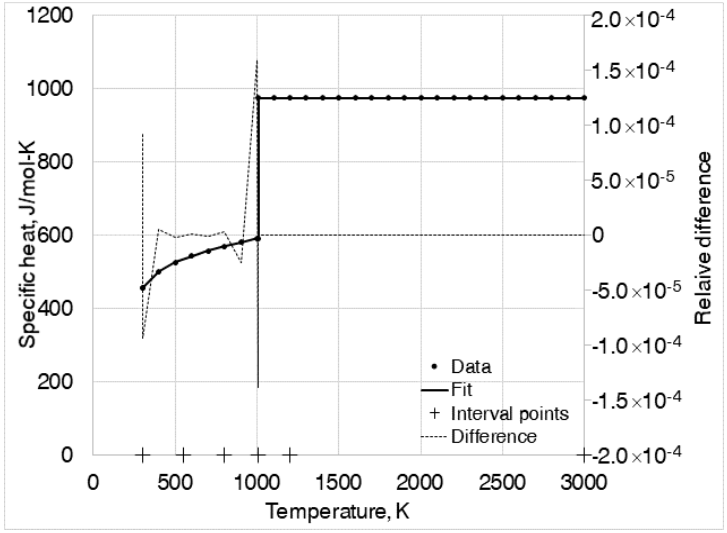


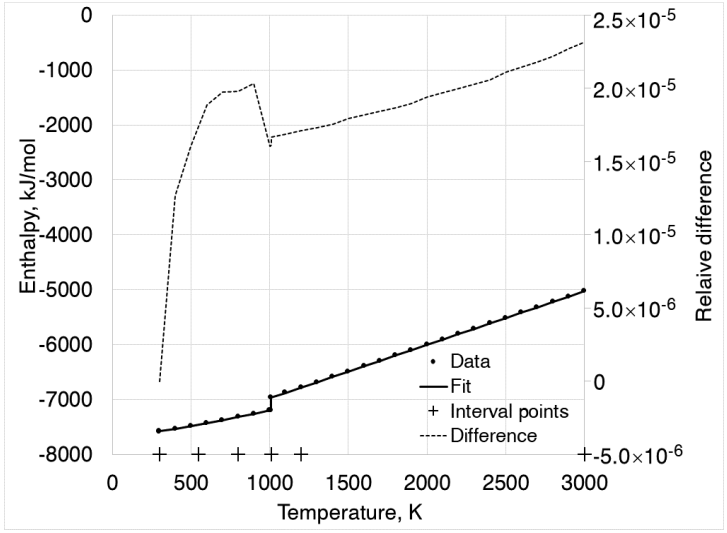


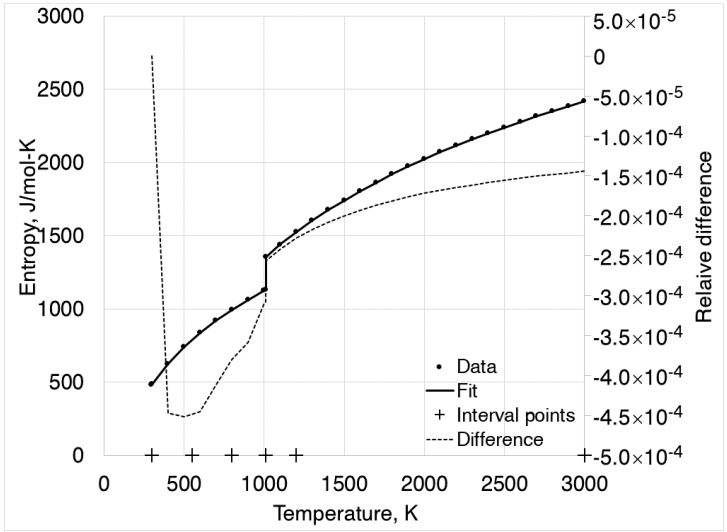


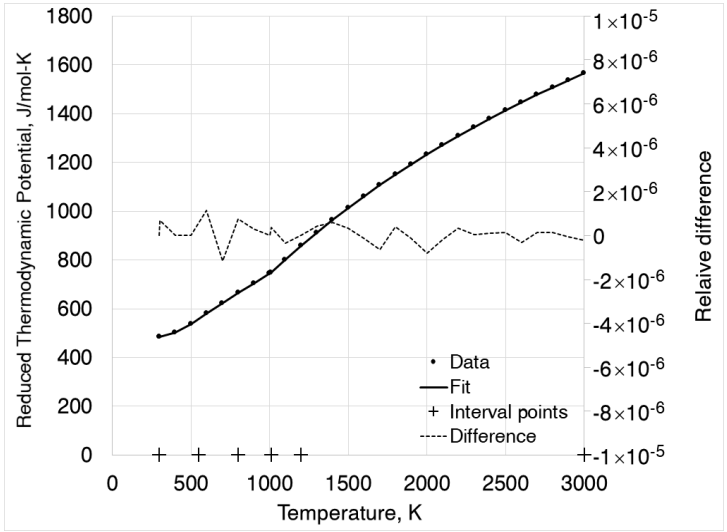
Chiolite ($Na_5Al_3F_{14}$) – Al3F14Na5(cr,l)

MW = 461.8711 g/mol – $\Delta_f H^o$(298.15 K) = -7581.408 kJ/mol – S^o(298.15 K) = 483.545 J/mol-K – p^o = 0.1 MPa



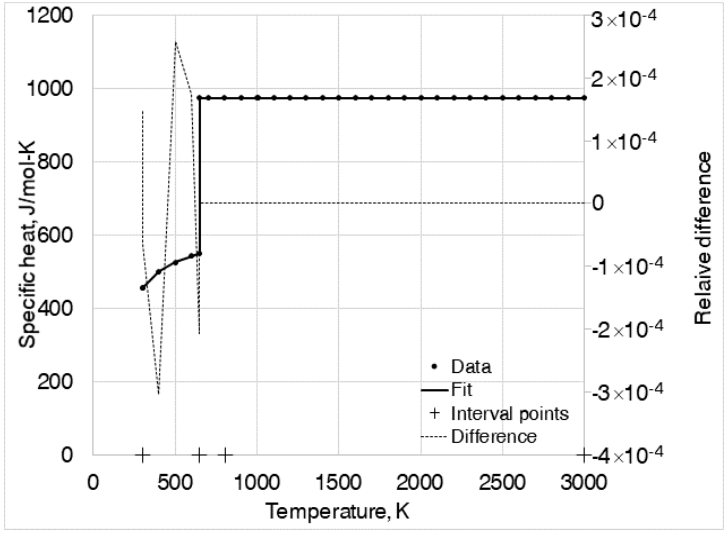


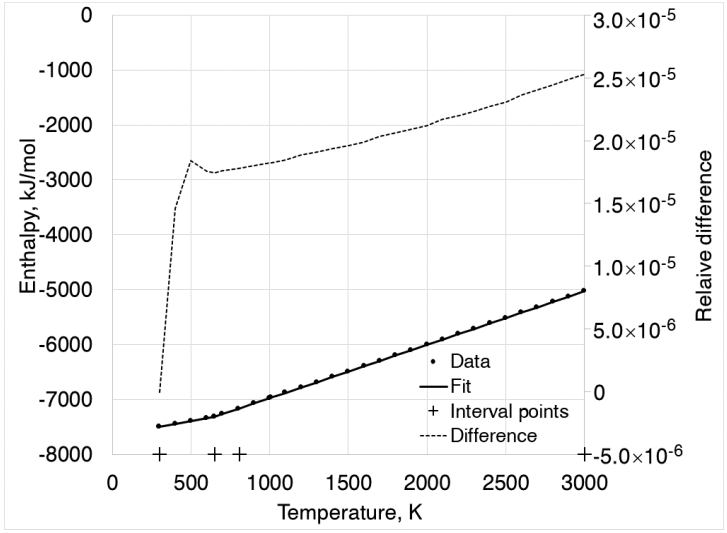


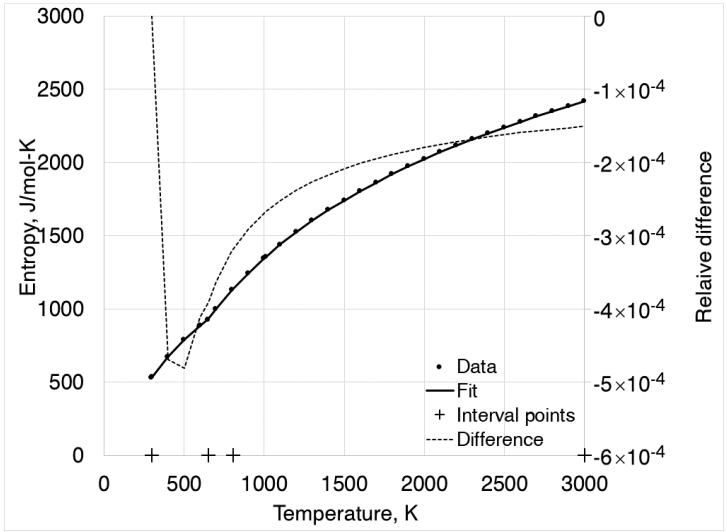


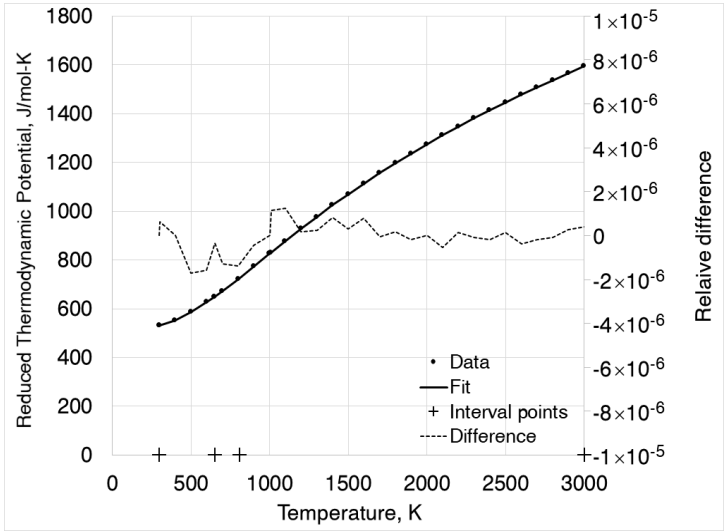
Chiolite ($Na_5Al_3F_{14}$) – Al3F14Na5(l)

MW = 461.8711 g/mol – $\Delta_f H^o$(298.15 K) = -7497.408 kJ/mol – S^o(298.15 K) = 532.124 J/mol-K – p^o = 0.1 MPa



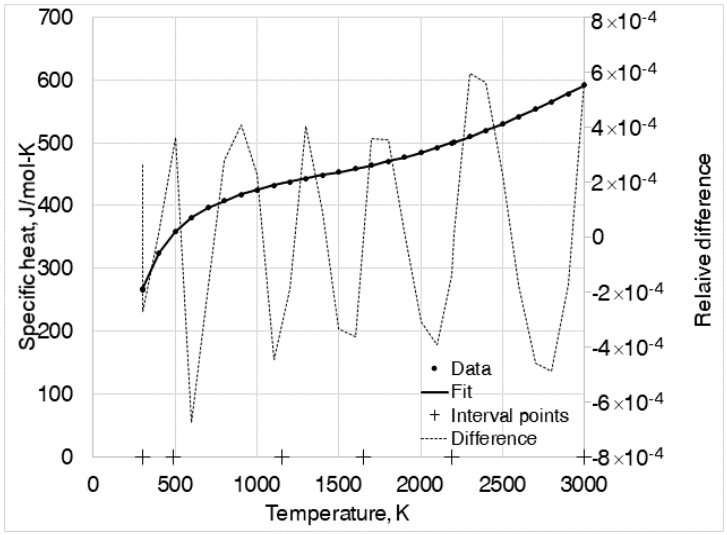


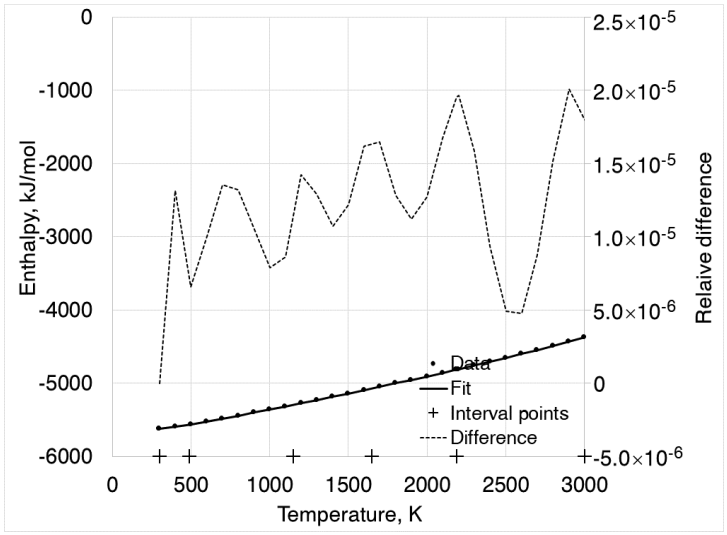


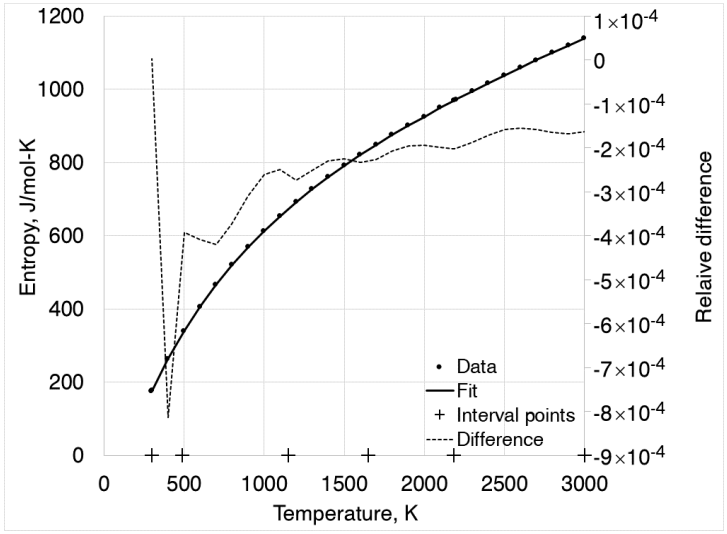


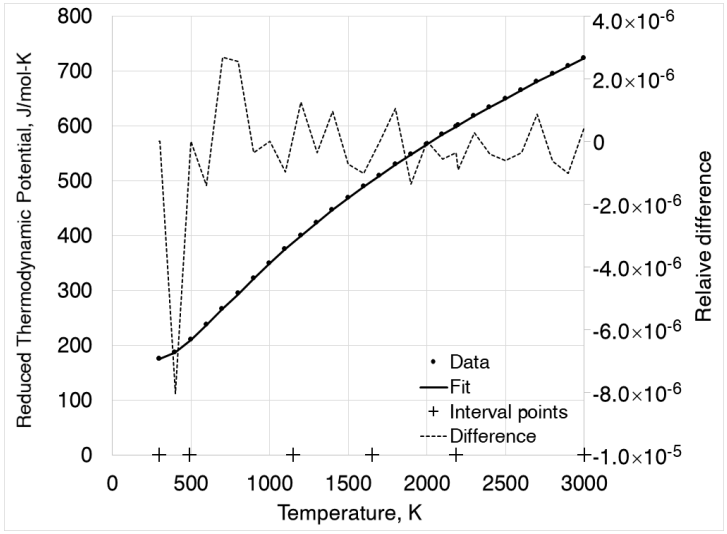
Beryllium Aluminum Oxide ($BeAl_6O_{10}$) – Al6Be1O10(cr)

MW = 330.8954 g/mol – $\Delta_f H^o$(298.15 K) = -5624.133 kJ/mol – S^o(298.15 K) = 175.552 J/mol-K – p^o = 0.1 MPa



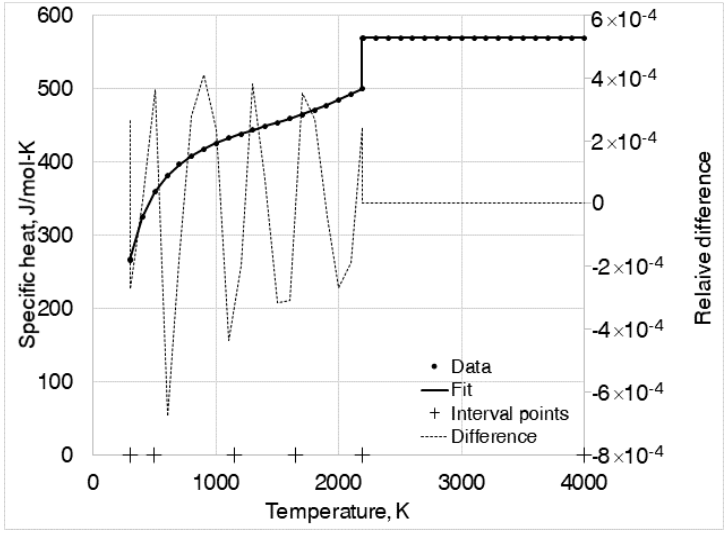


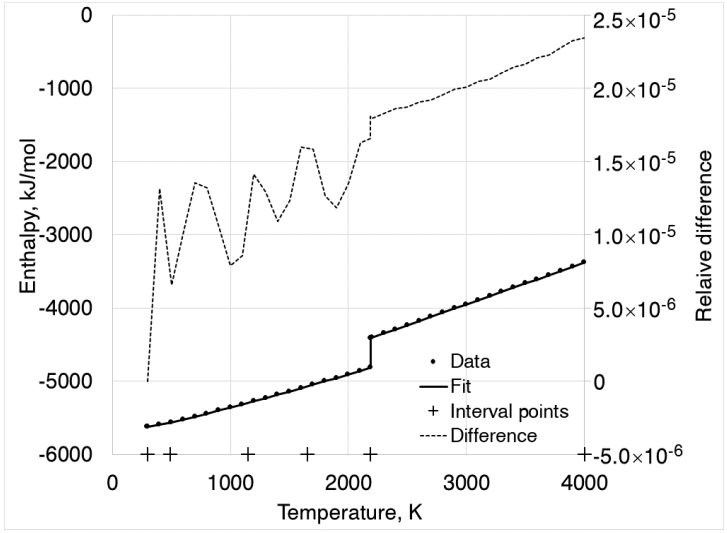


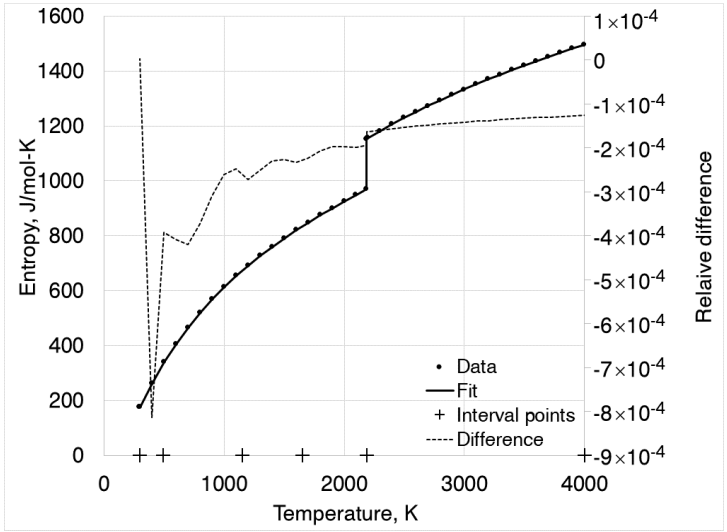


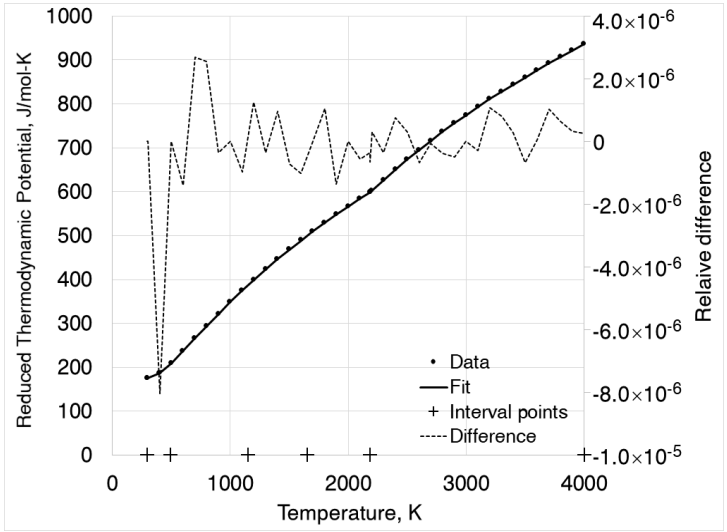
Beryllium Aluminum Oxide ($BeAl_6O_{10}$) – Al6Be1O10(cr,l)

MW = 330.8954 g/mol – $\Delta_f H^o$(298.15 K) = -5624.133 kJ/mol – S^o(298.15 K) = 175.552 J/mol-K – p^o = 0.1 MPa



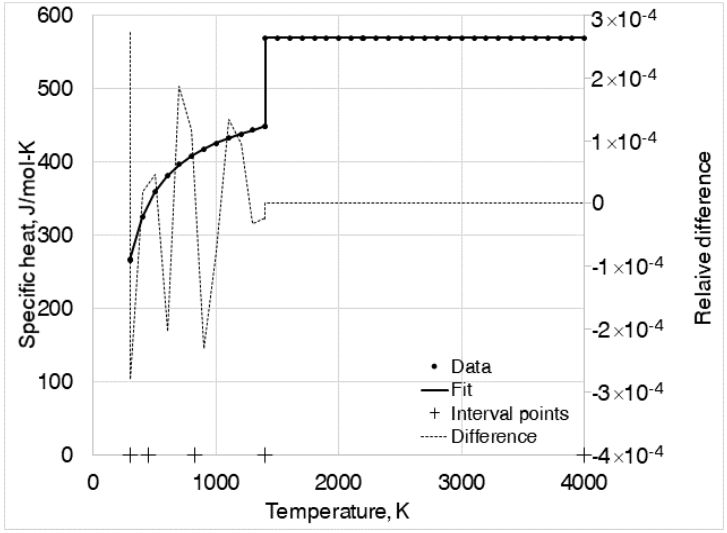


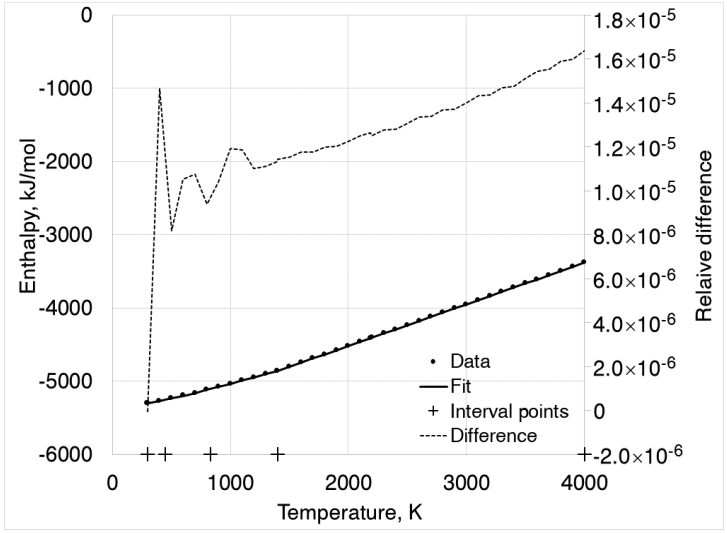


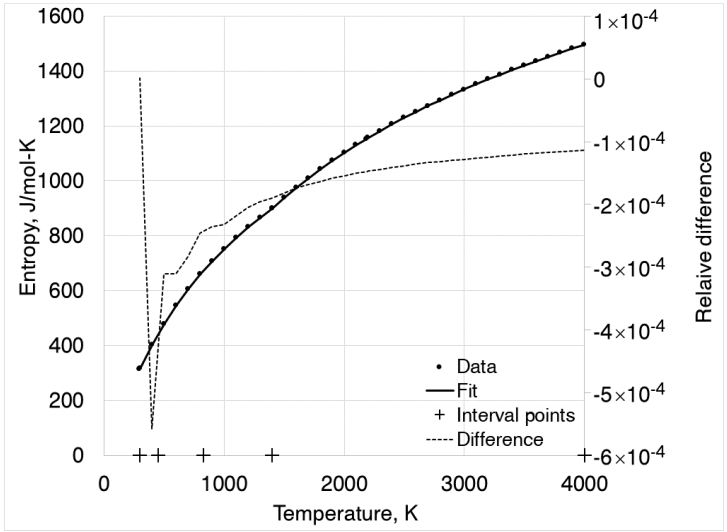


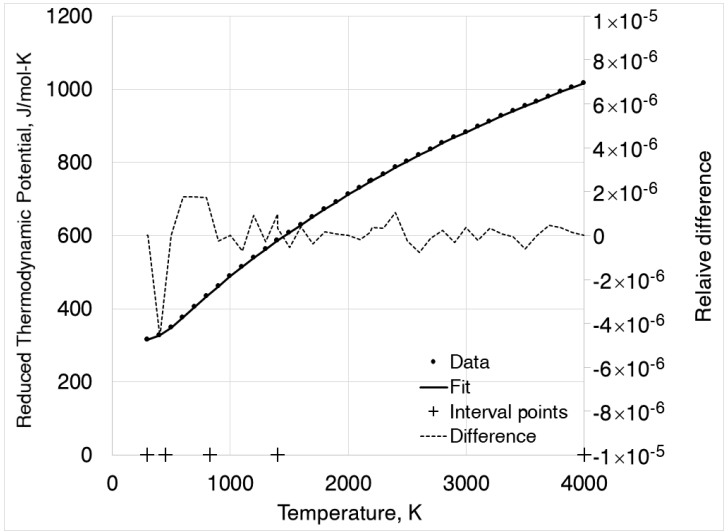
Beryllium Aluminum Oxide ($BeAl_6O_{10}$) – Al6Be1O10(l)

MW = 330.8954 g/mol – $\Delta_f H^o$(298.15 K) = -5299.312 kJ/mol – S^o(298.15 K) = 314.903 J/mol-K – p^o = 0.1 MPa



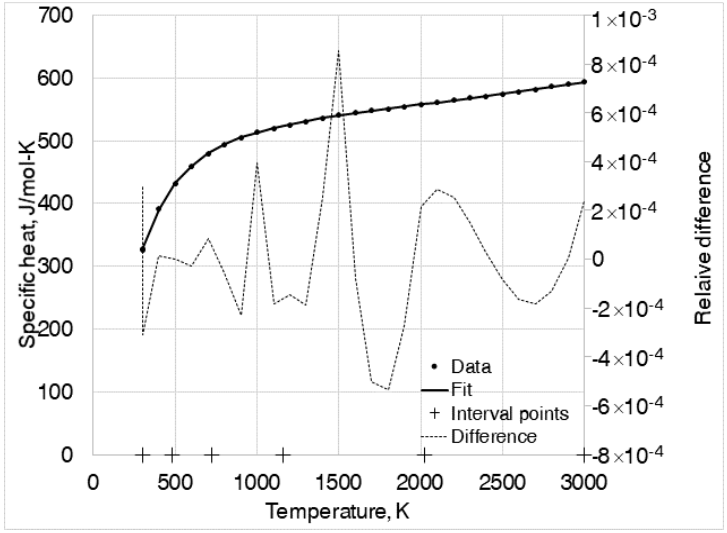


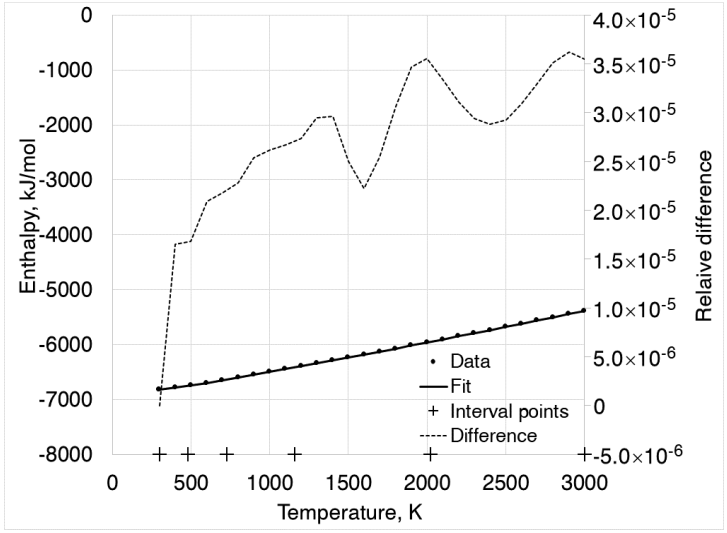


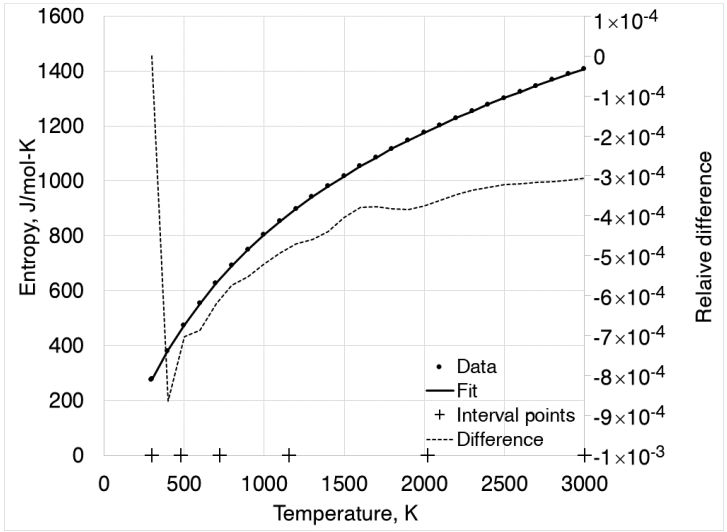


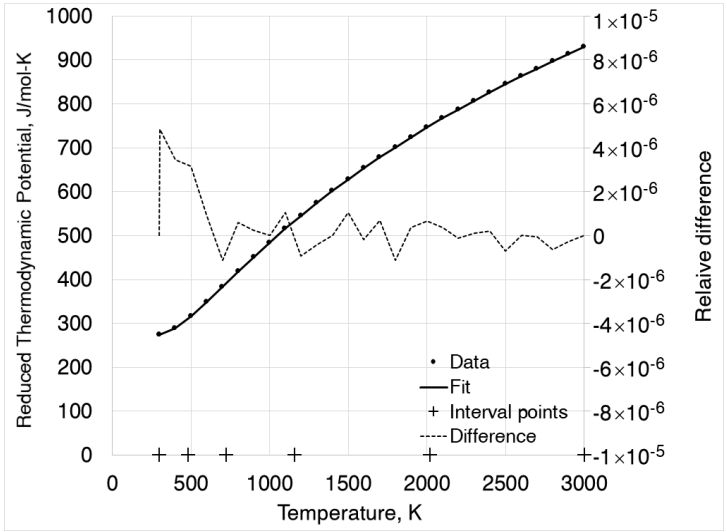
Aluminum Silicate, Mullite ($Al_6Si_2O_{13}$) – Al6O13Si2(cr)

MW = 426.0524 g/mol – $\Delta_f H^o$(298.15 K) = -6819.209 kJ/mol – S^o(298.15 K) = 274.889 J/mol-K – p^o = 0.1 MPa



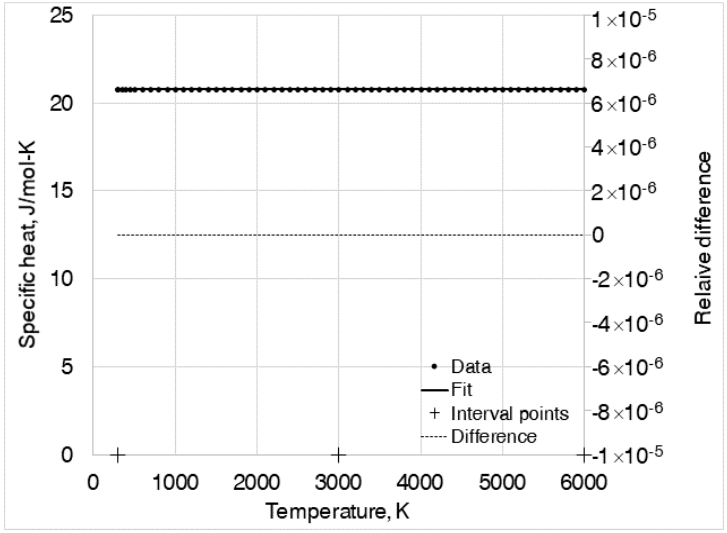


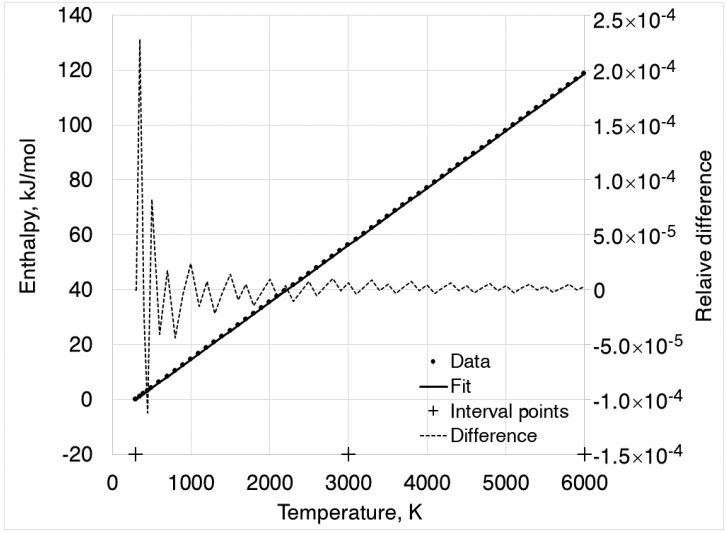


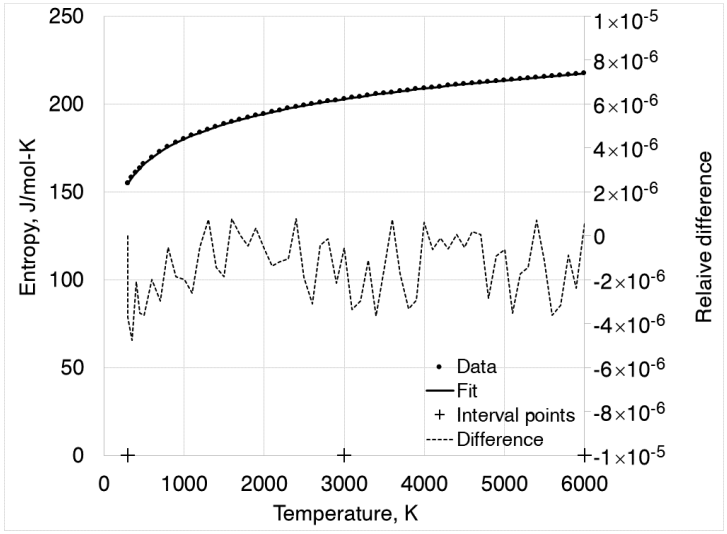


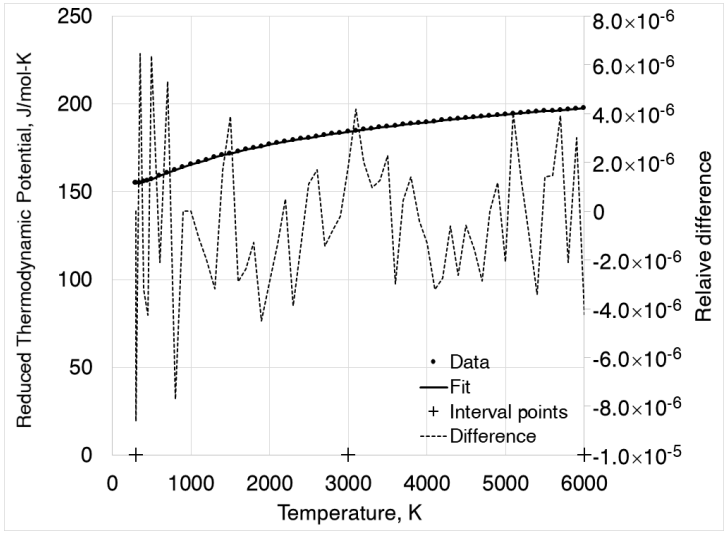
Argon (Ar) – Ar1(ref)

MW = 39.948 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 154.845 J/mol-K – p^o = 0.1 MPa



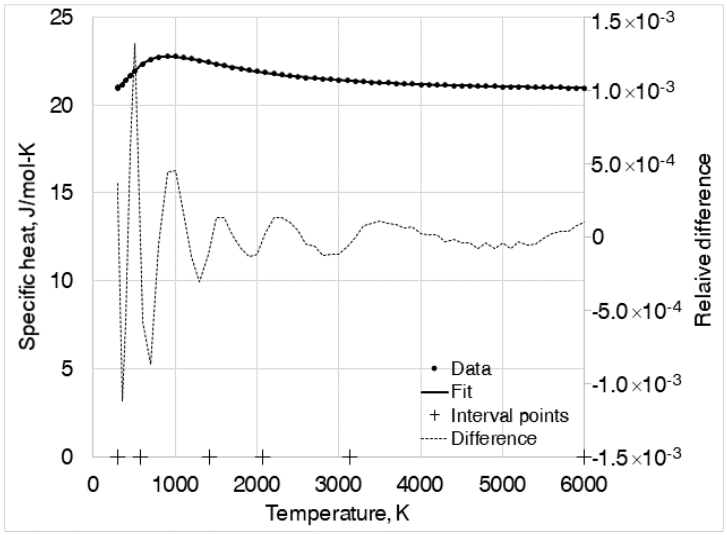


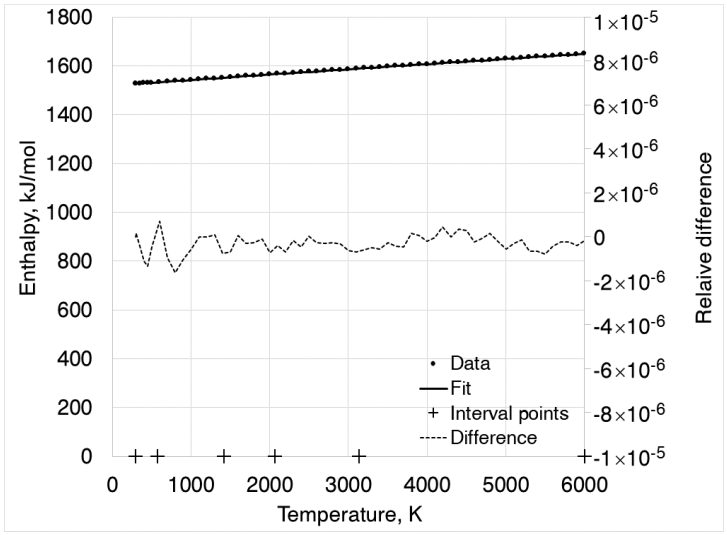


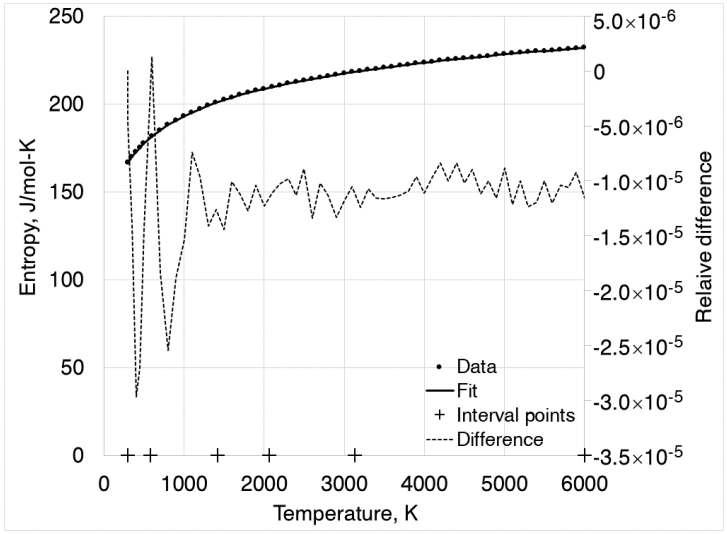


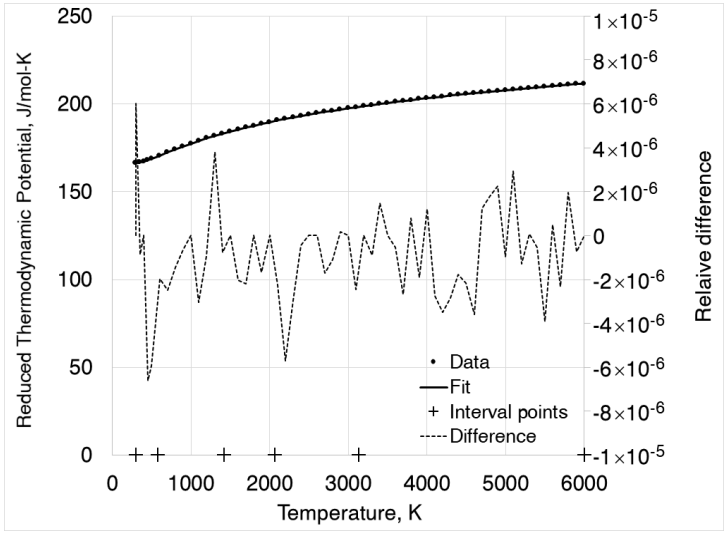
Argon, Ion (Ar^+) – Ar1+(g)

MW = 39.947 g/mol – $\Delta_f H^o$(298.15 K) = 1526.778 kJ/mol – S^o(298.15 K) = 166.404 J/mol-K – p^o = 0.1 MPa



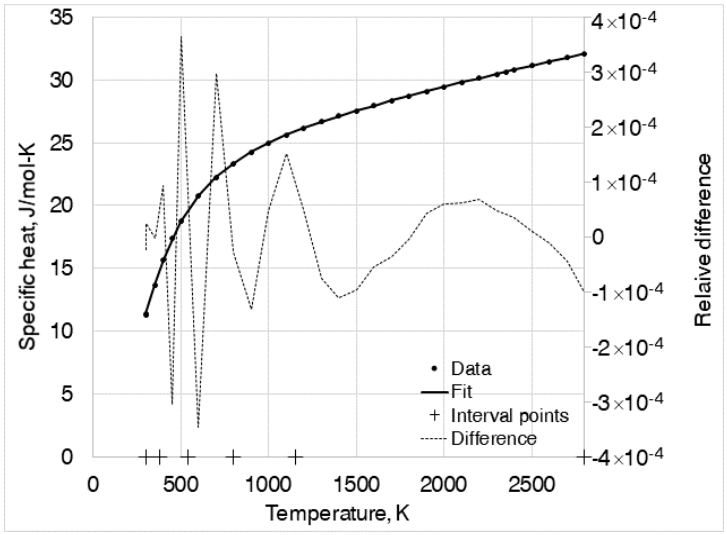


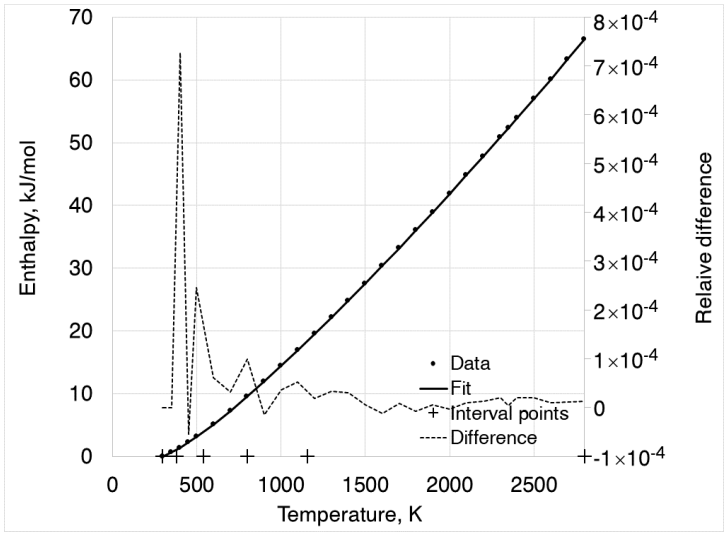


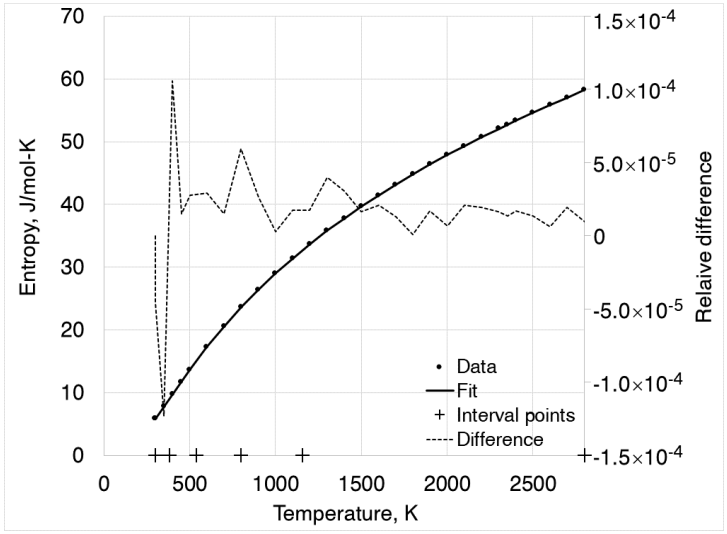


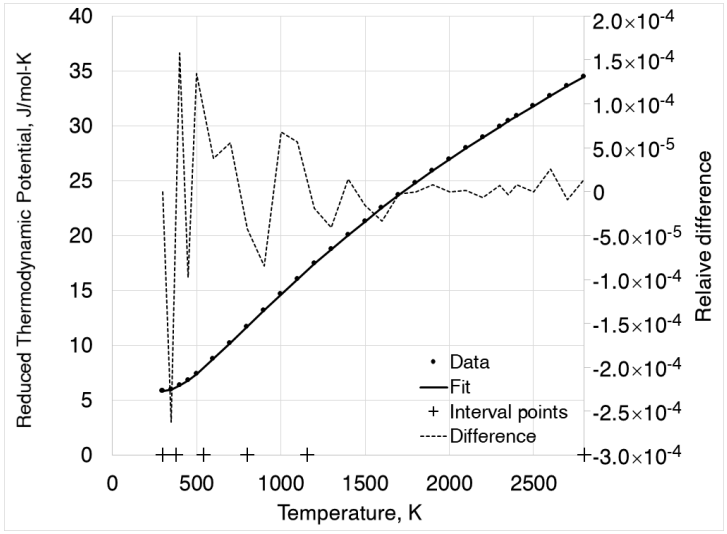
Boron, Beta-Rhombohedral (B) – B1(cr)

MW = 10.81 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 5.834 J/mol-K – p^o = 0.1 MPa



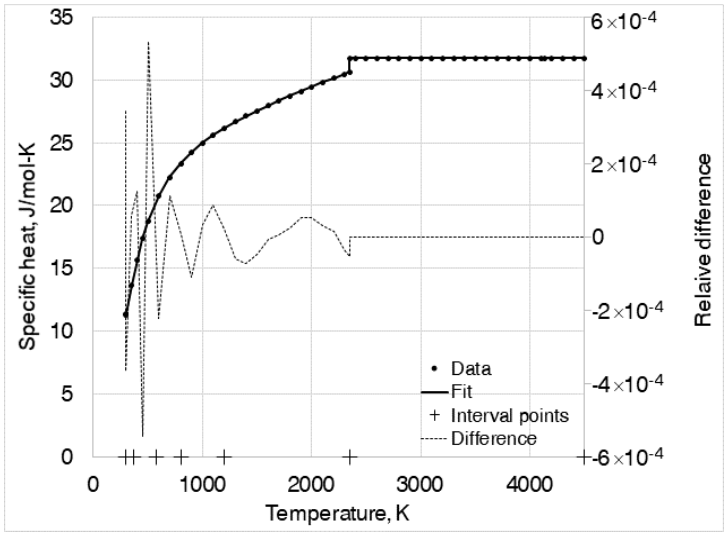


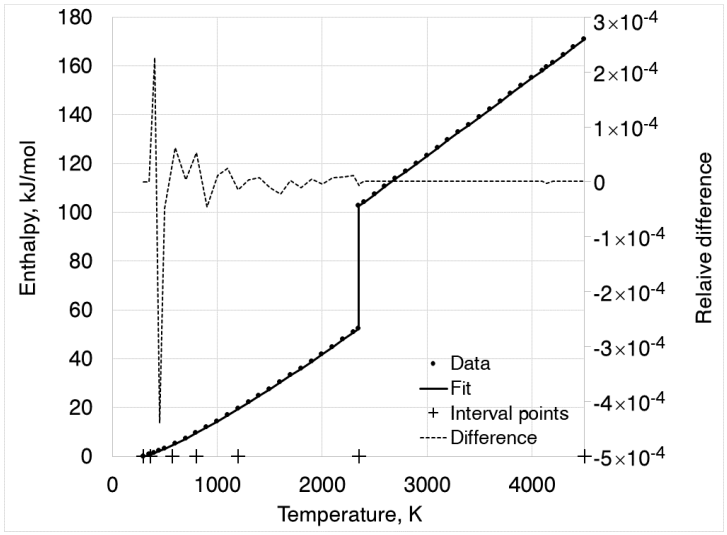


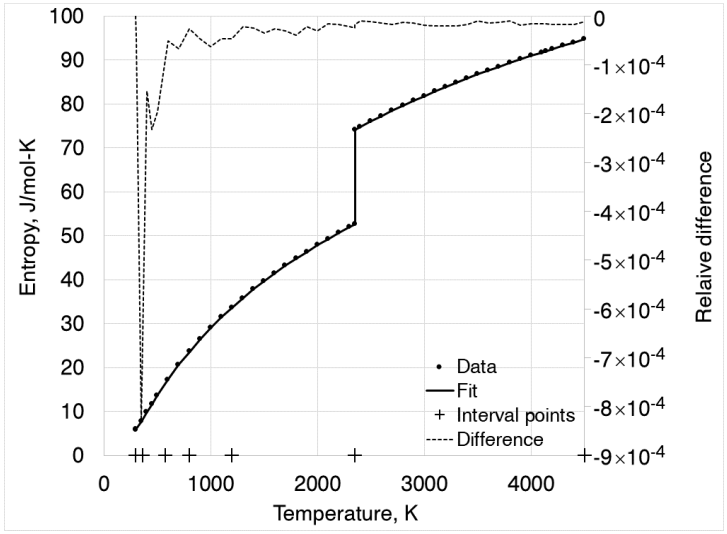


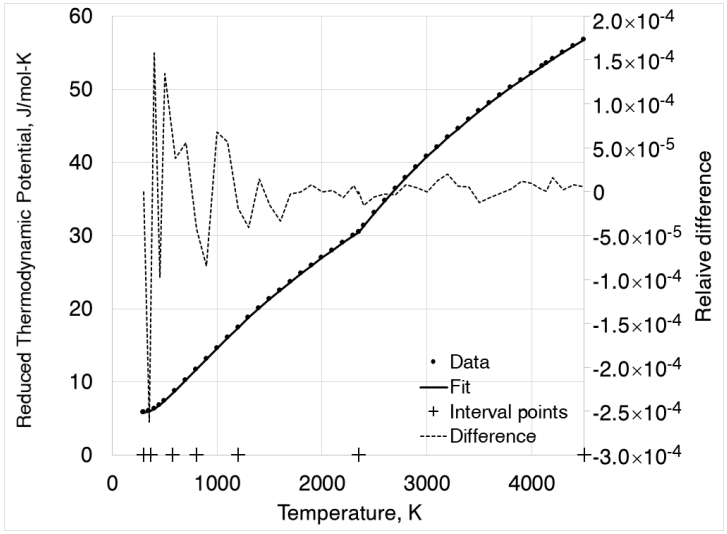
Boron (B) – B1(cr,l)

MW = 10.81 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 5.834 J/mol-K – p^o = 0.1 MPa



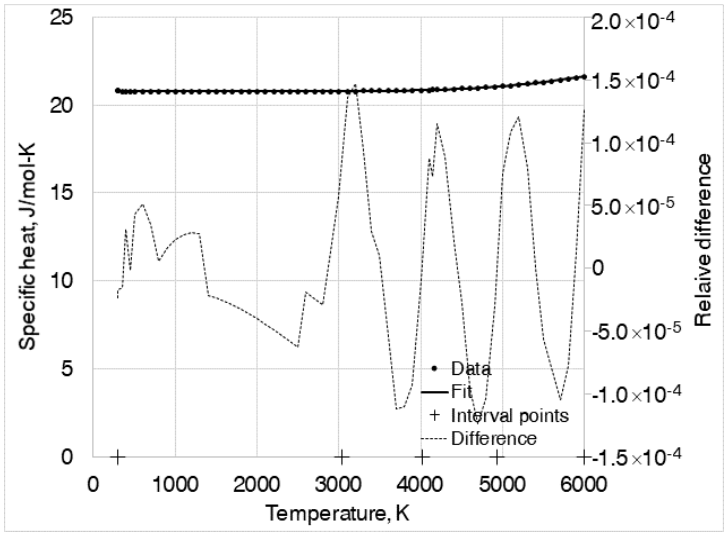


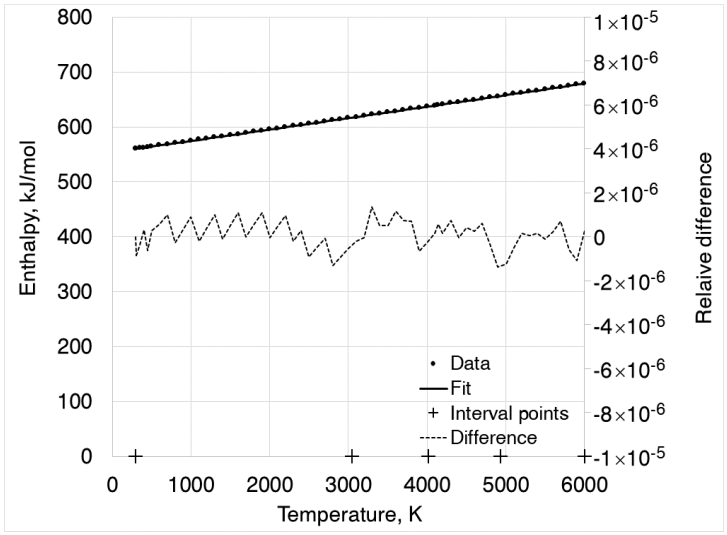


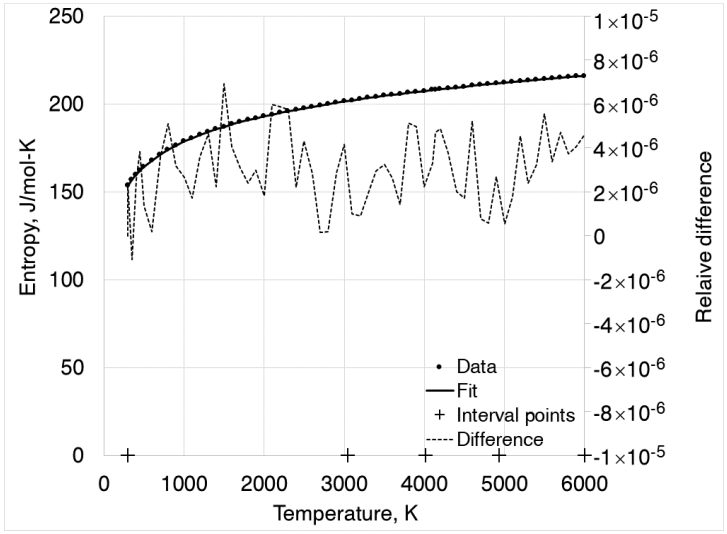


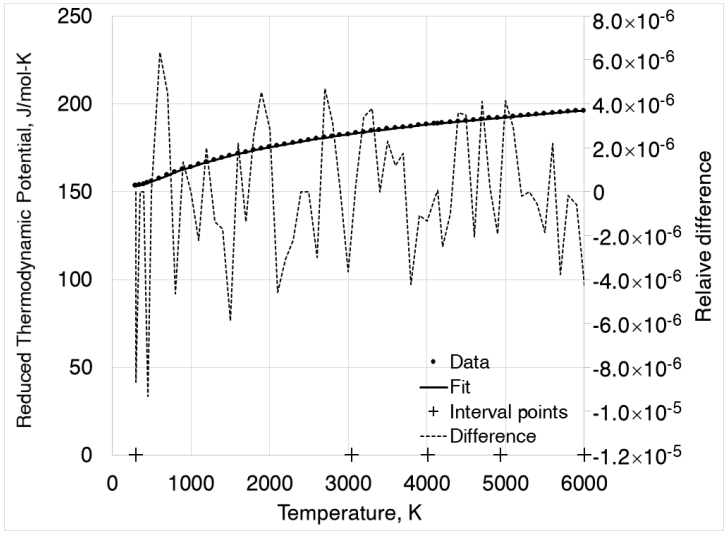
Boron (B) – B1(g)

MW = 10.81 g/mol – $\Delta_f H^o$(298.15 K) = 560 kJ/mol – S^o(298.15 K) = 153.435 J/mol-K – p^o = 0.1 MPa



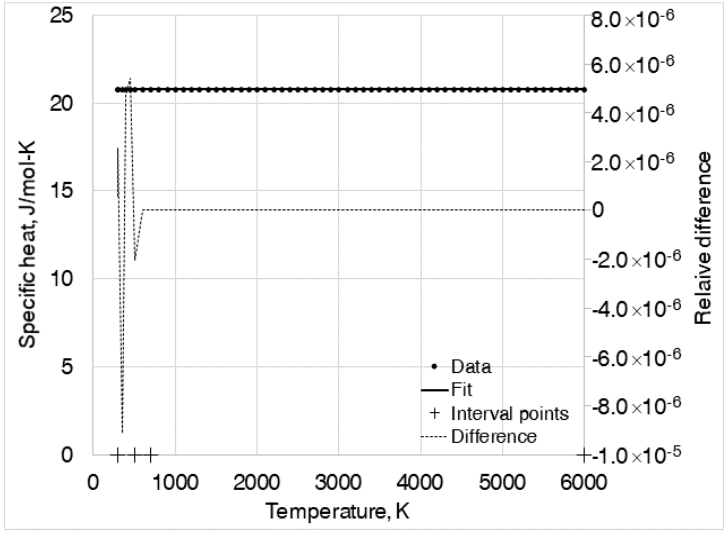


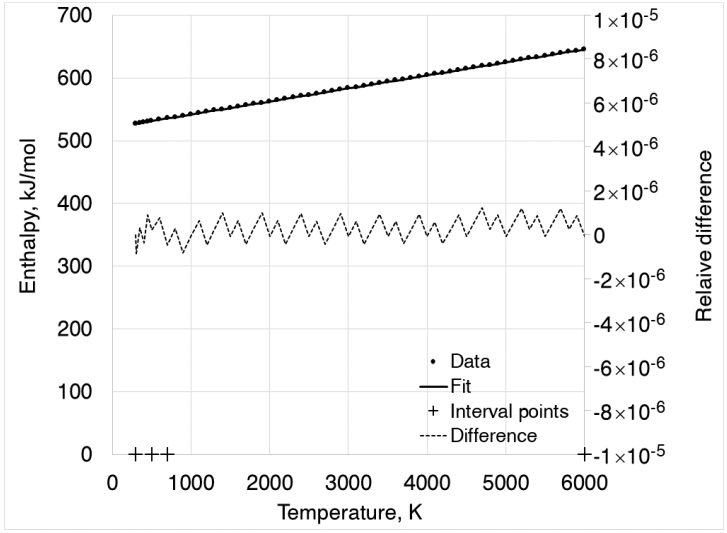


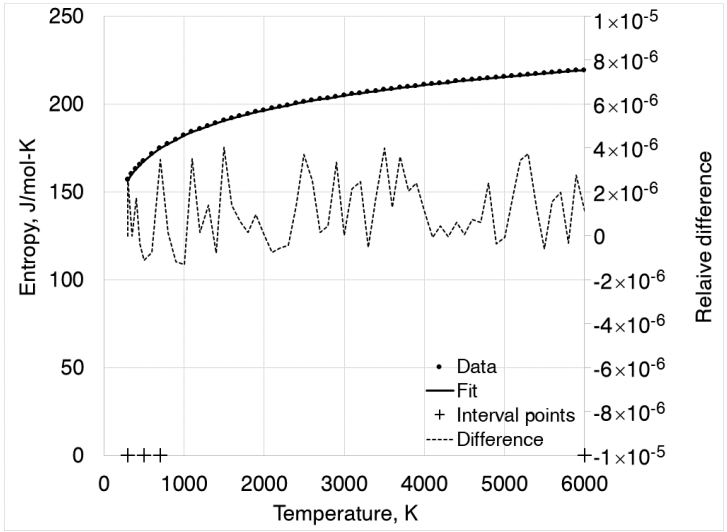


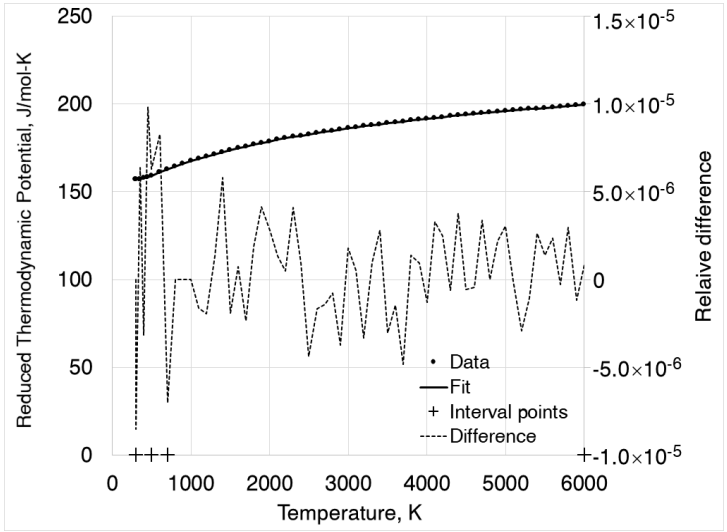
Boron, Ion (B^-) – B1-(g)

MW = 10.81 g/mol – $\Delta_f H^o$(298.15 K) = 527.033 kJ/mol – S^o(298.15 K) = 156.811 J/mol-K – p^o = 0.1 MPa



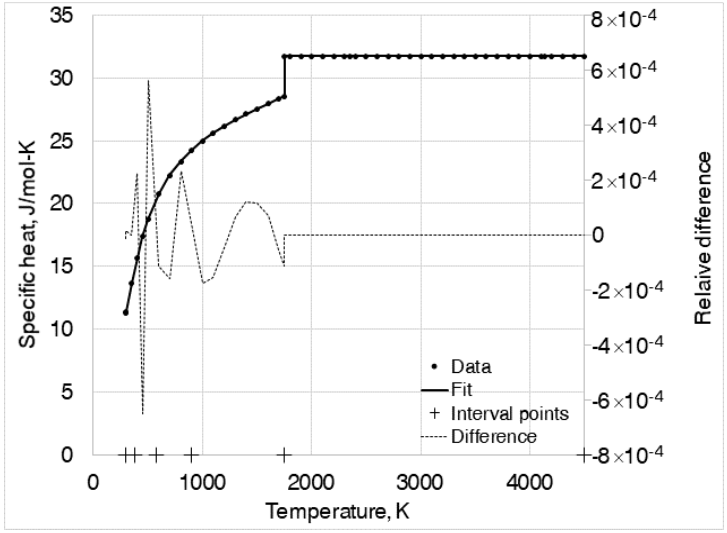


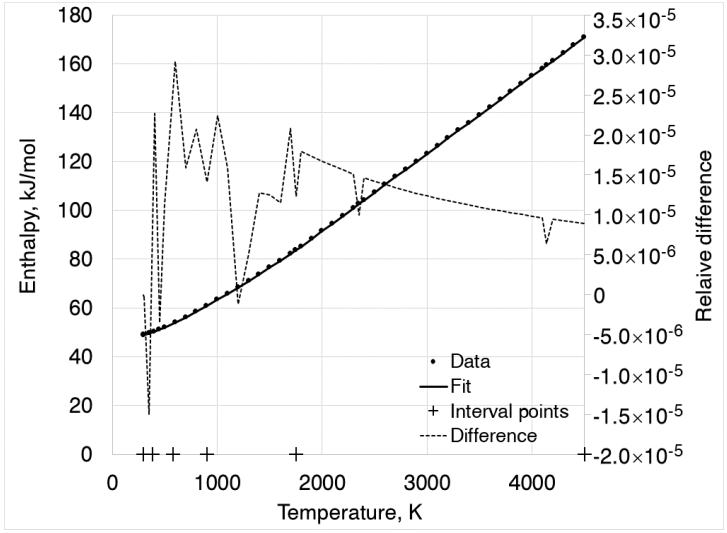


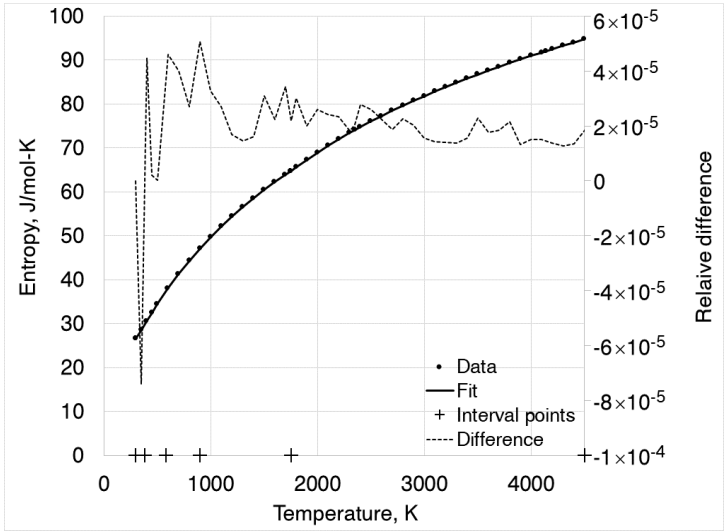


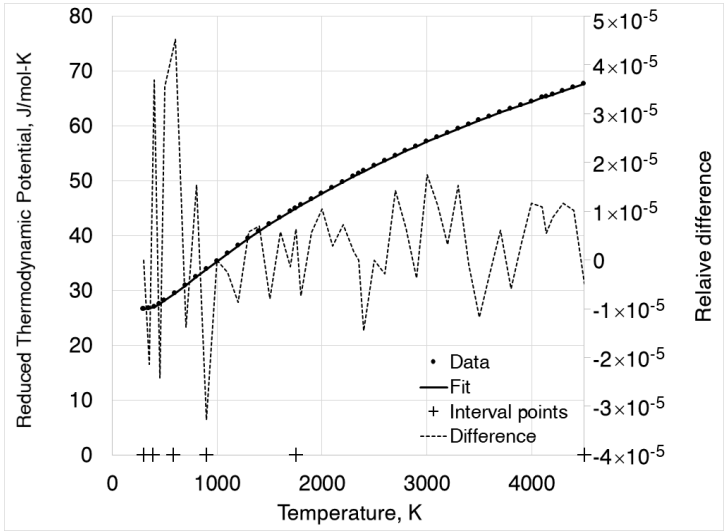
Boron (B) – B1(l)

MW = 10.81 g/mol – $\Delta_f H^o$(298.15 K) = 48.927 kJ/mol – S^o(298.15 K) = 26.555 J/mol-K – p^o = 0.1 MPa



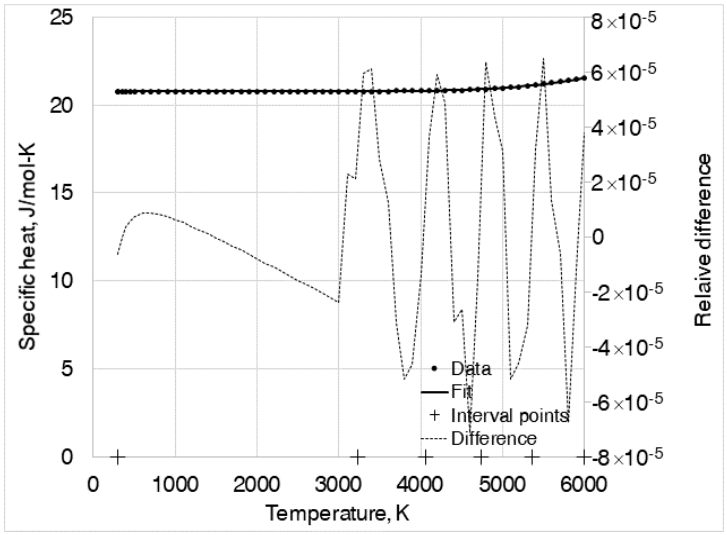


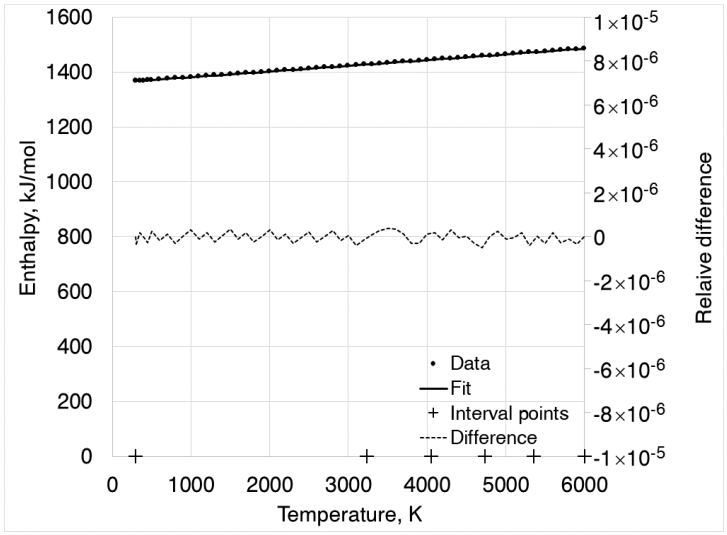


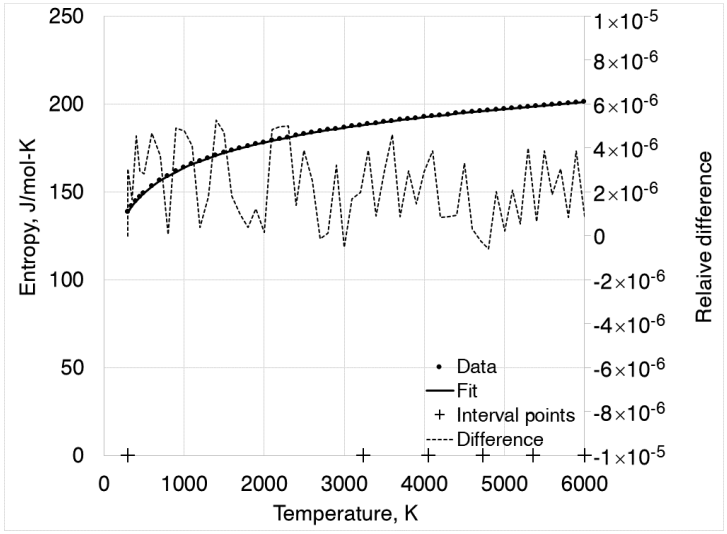


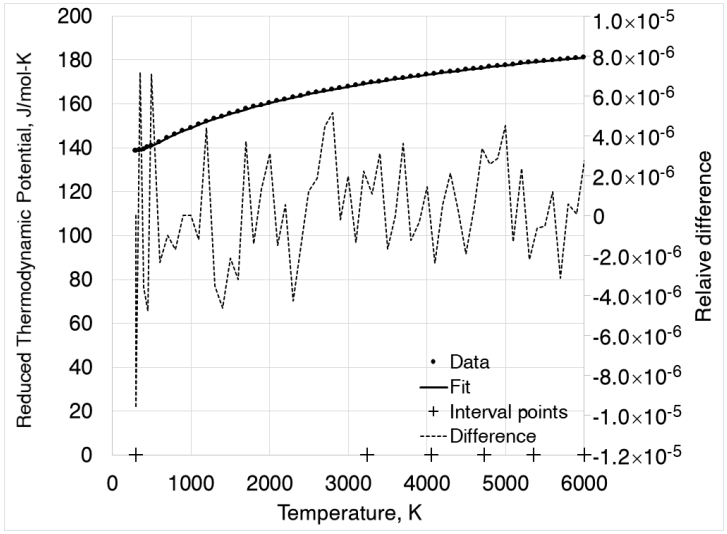
Boron, Ion (B^+) – B1+(g)

MW = 10.81 g/mol – $\Delta_f H^o$(298.15 K) = 1366.717 kJ/mol – S^o(298.15 K) = 138.542 J/mol-K – p^o = 0.1 MPa



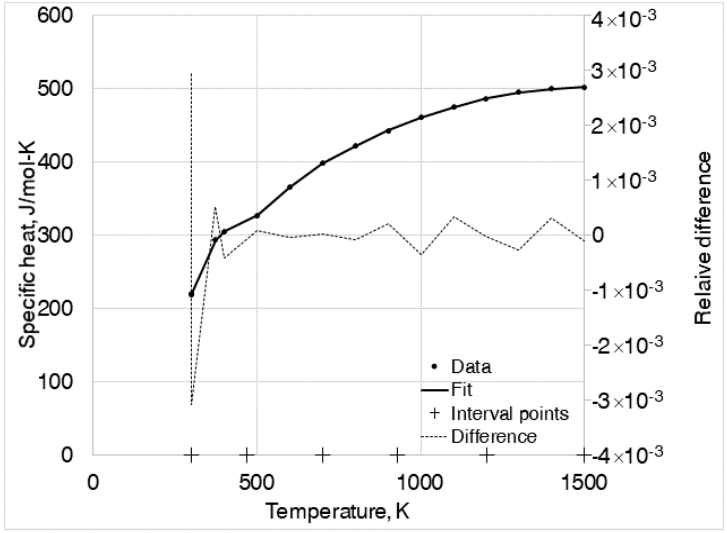


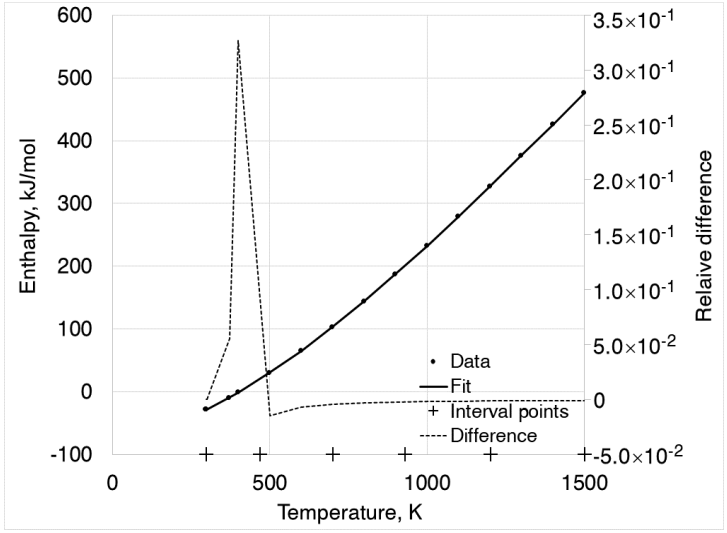


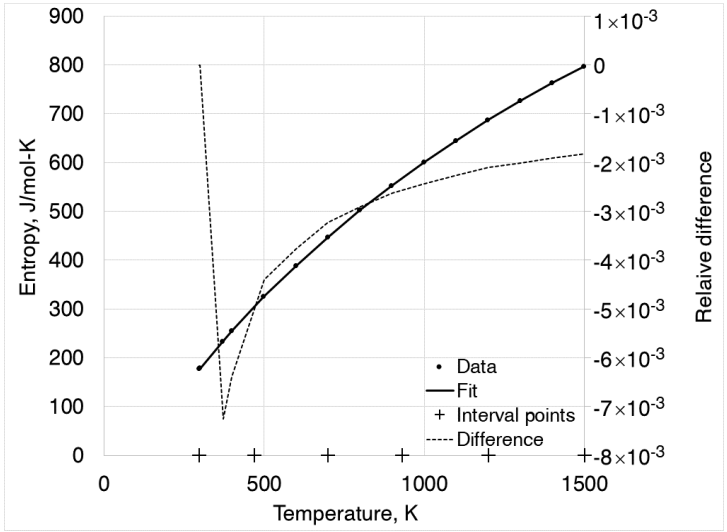


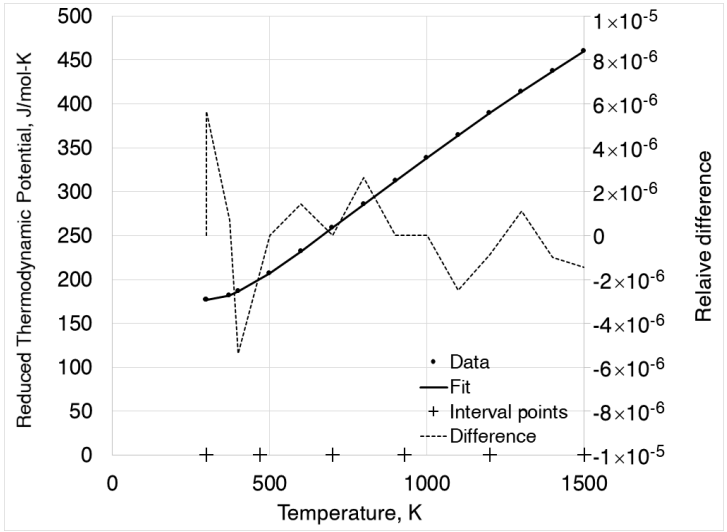
Decaborane ($B_{10}H_{14}$) – B10H14(cr)

MW = 122.21 g/mol – $\Delta_f H^o$(298.15 K) = -28.87 kJ/mol – S^o(298.15 K) = 176.565 J/mol-K – p^o = 0.1 MPa



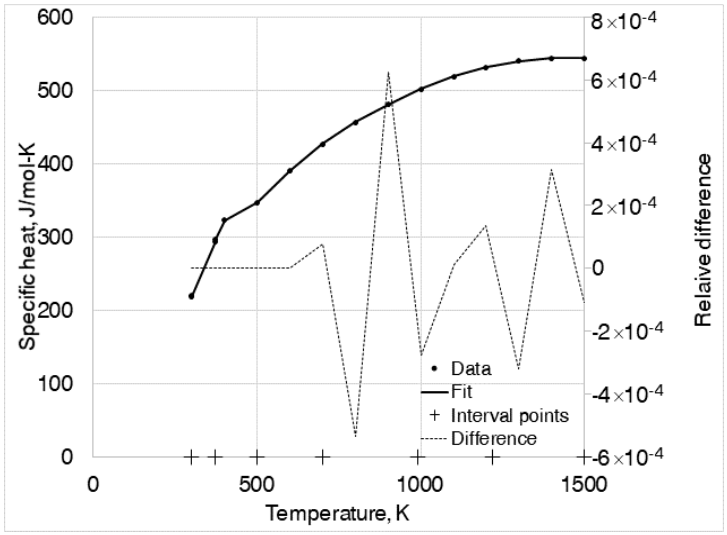


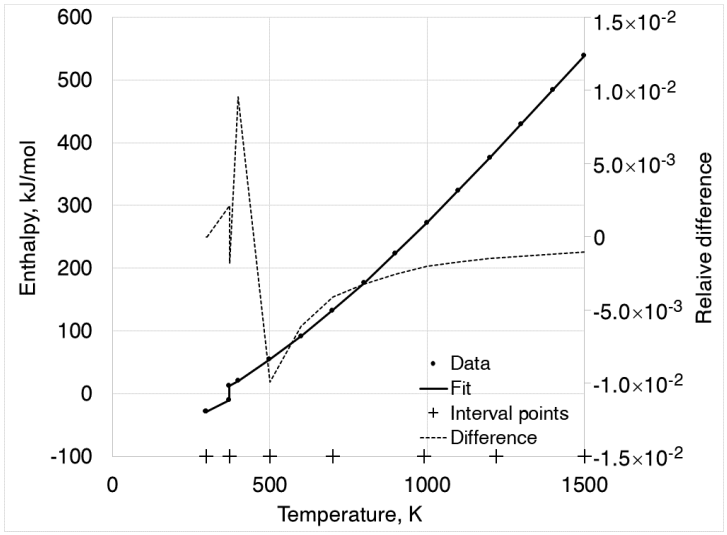


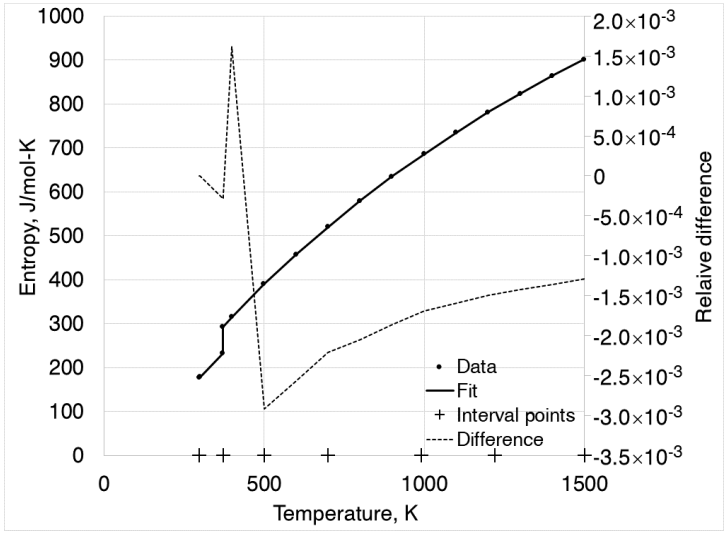


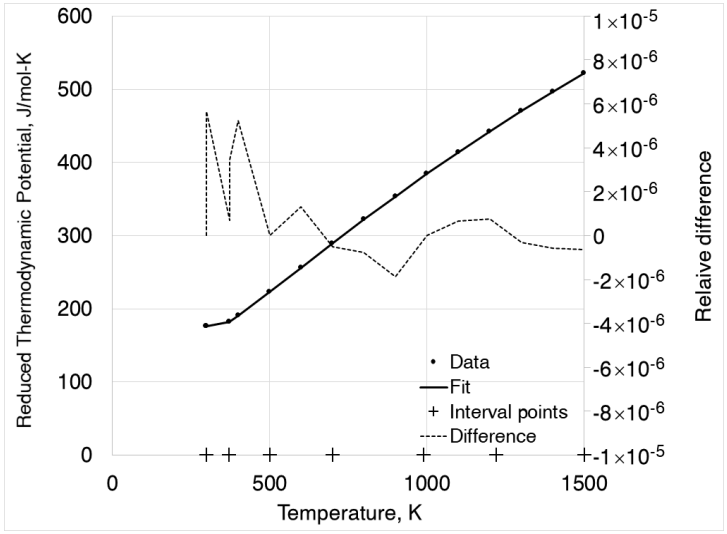
Decaborane ($B_{10}H_{14}$) – B10H14(cr,l)

MW = 122.21 g/mol – $\Delta_f H^o$(298.15 K) = -28.87 kJ/mol – S^o(298.15 K) = 176.565 J/mol-K – p^o = 0.1 MPa



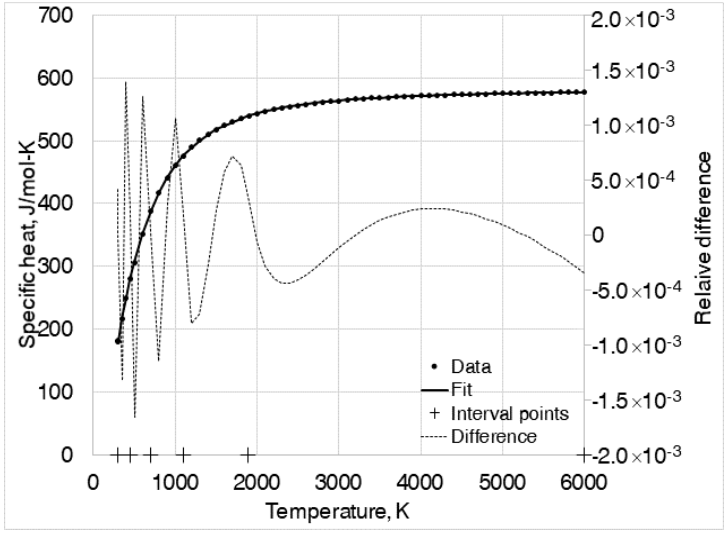


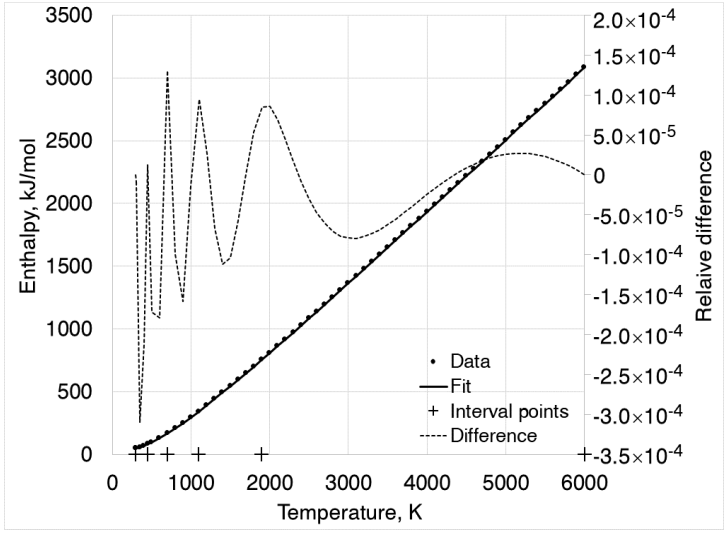


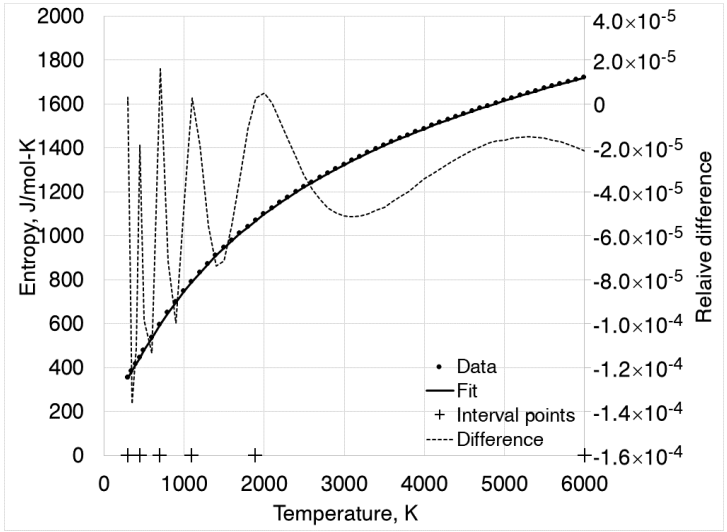


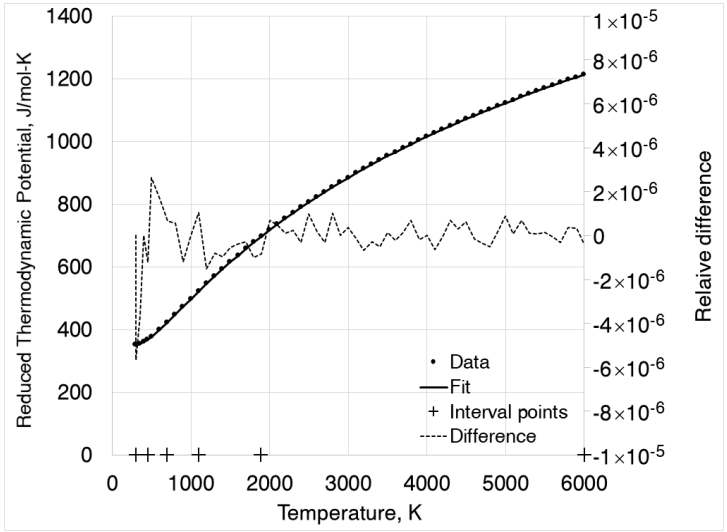
Decaborane ($B_{10}H_{14}$) – B10H14(g)

MW = 122.21 g/mol – $\Delta_f H^o$(298.15 K) = 47.279 kJ/mol – S^o(298.15 K) = 352.159 J/mol-K – p^o = 0.1 MPa



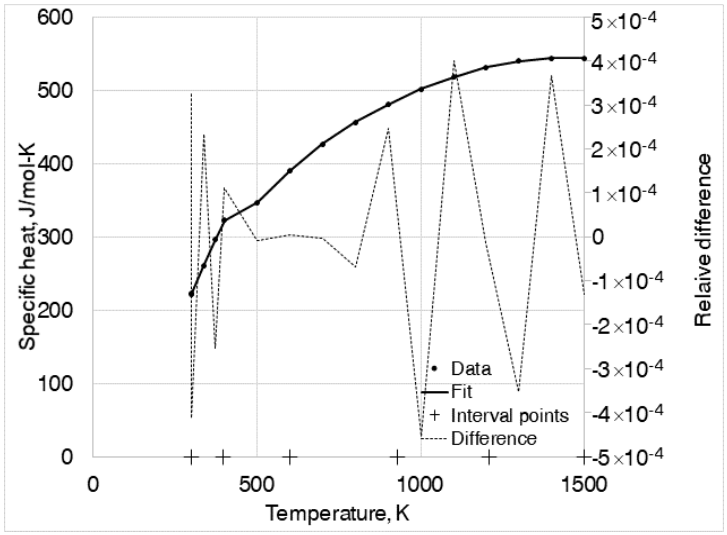


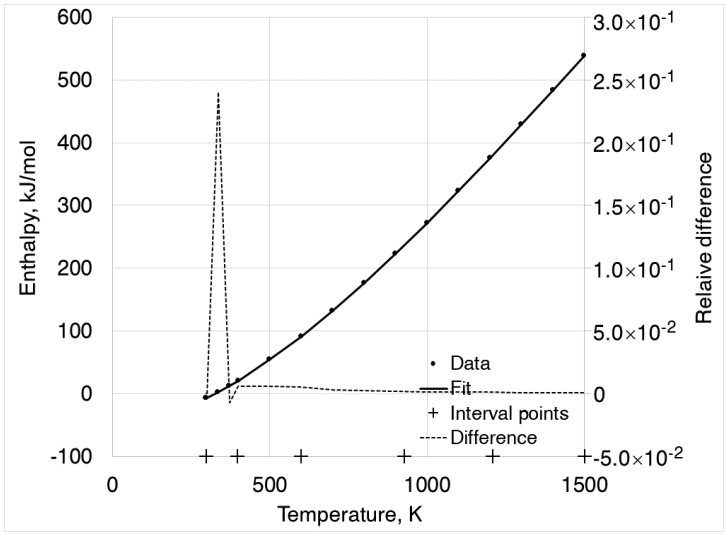


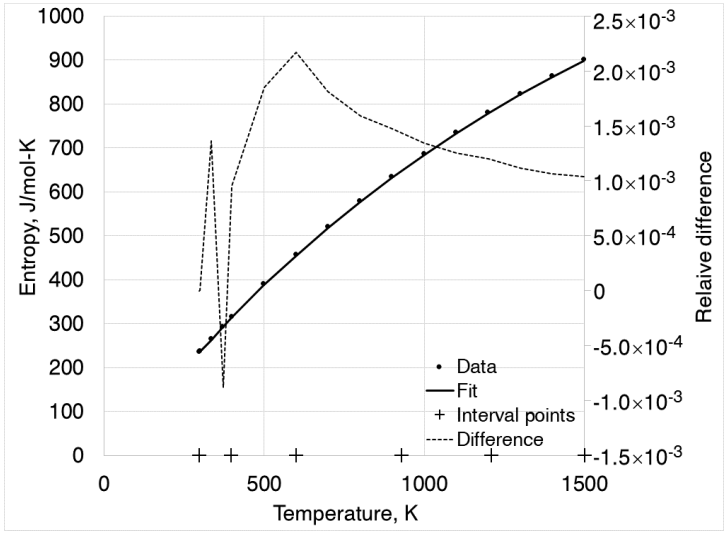


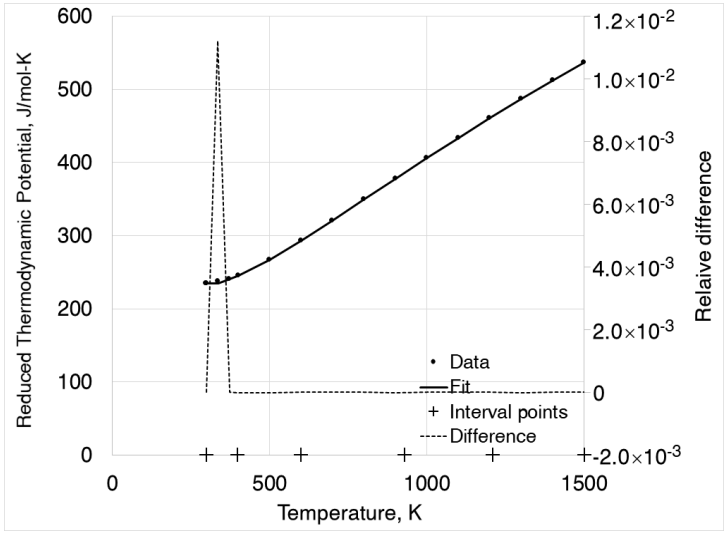
Decaborane ($B_{10}H_{14}$) – B10H14(l)

MW = 122.21 g/mol – $\Delta_f H^o$(298.15 K) = -7.133 kJ/mol – S^o(298.15 K) = 234.935 J/mol-K – p^o = 0.1 MPa



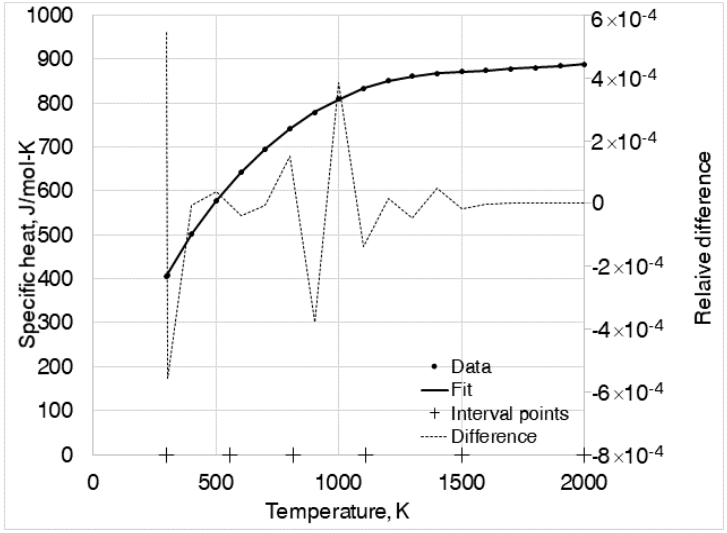


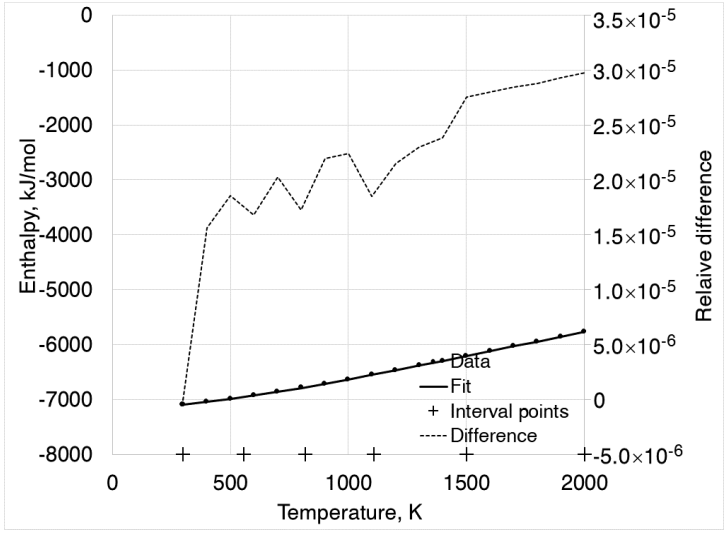


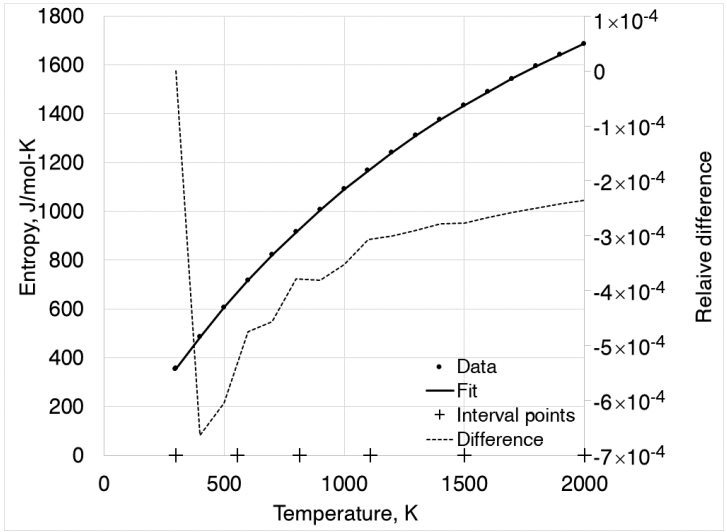


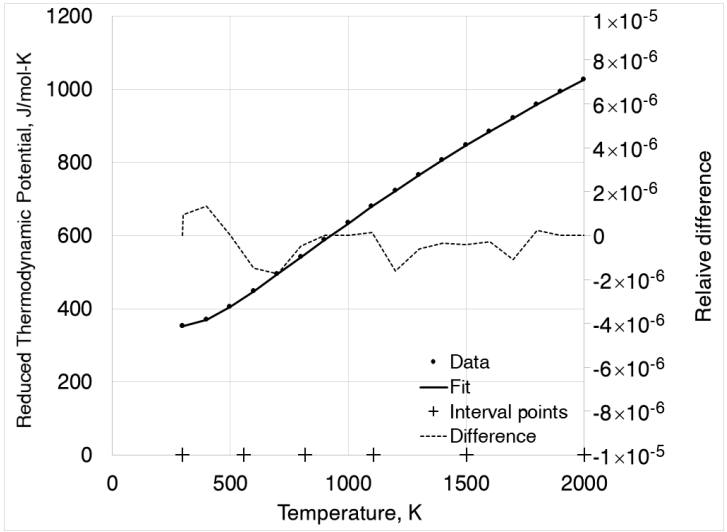
Lead Borate ($Pb_2B_{10}O_{17}$) – B10O17Pb2(cr)

MW = 794.5 g/mol – $\Delta_f H^o$(298.15 K) = -7087.696 kJ/mol – S^o(298.15 K) = 352.711 J/mol-K – p^o = 0.1 MPa



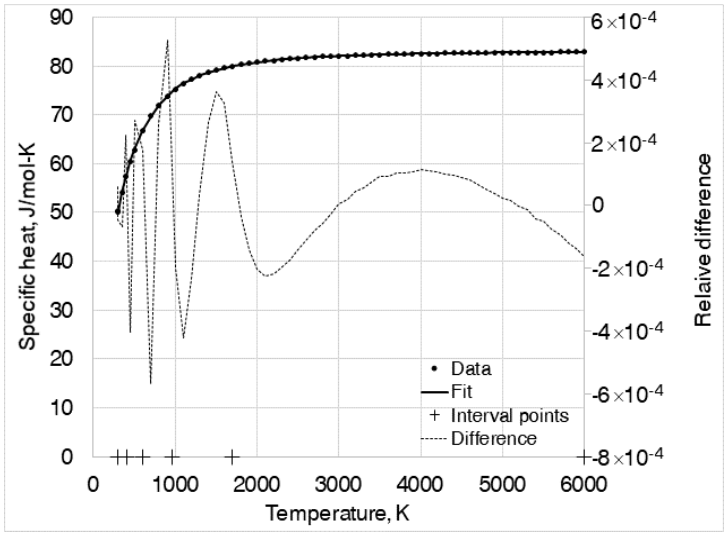


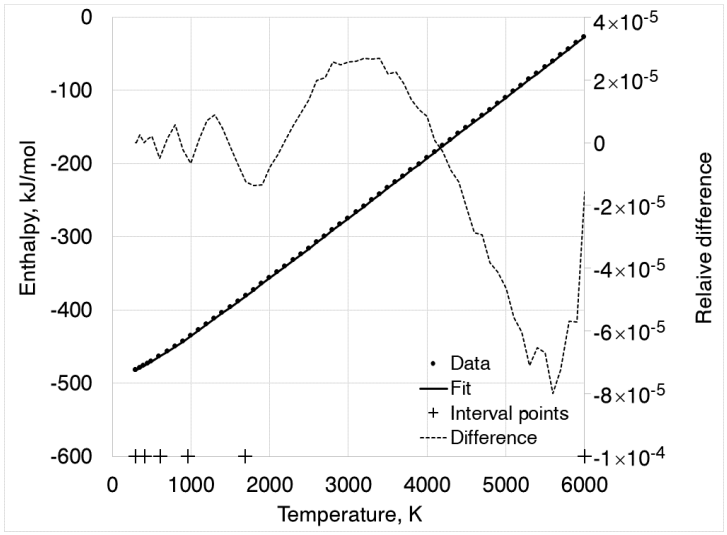


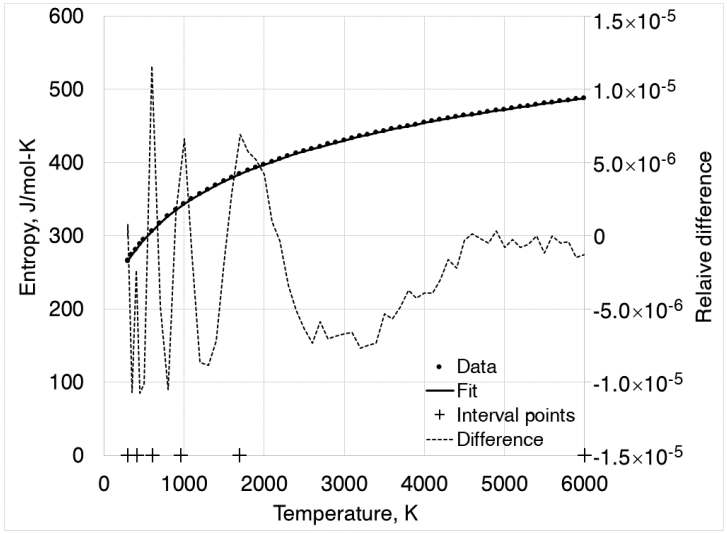


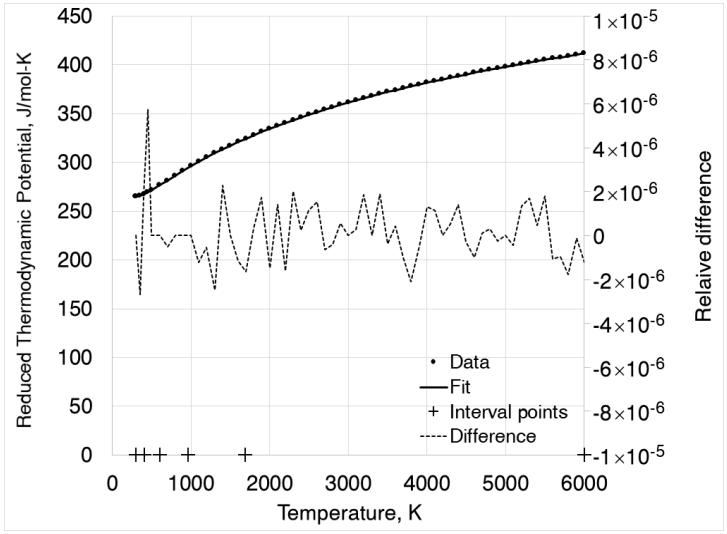
Beryllium Borate ($BeBO_2$) – B1Be1O2(g)

MW = 51.82 g/mol – $\Delta_f H^o$(298.15 K) = -481.997 kJ/mol – S^o(298.15 K) = 265.396 J/mol-K – p^o = 0.1 MPa



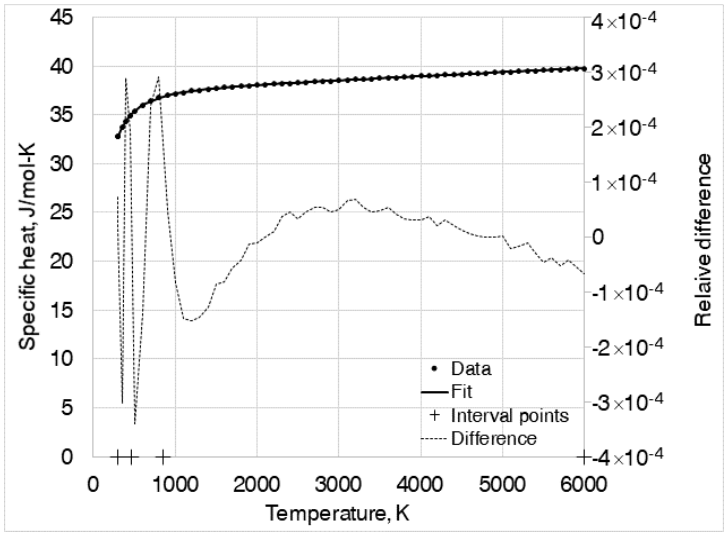


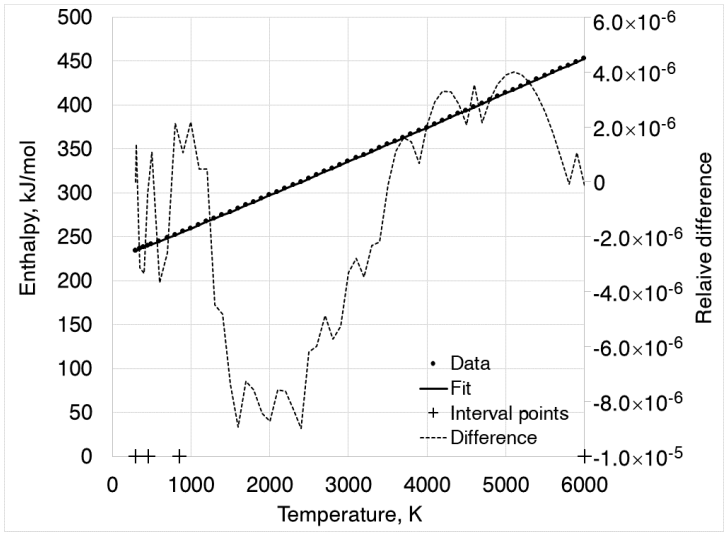


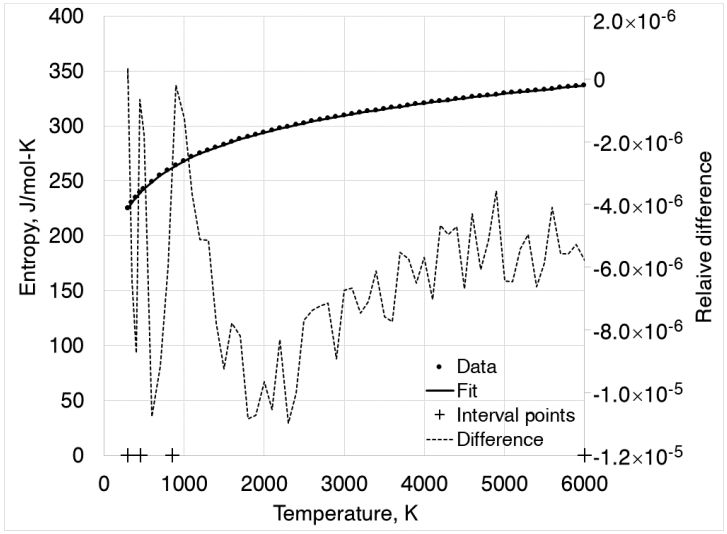


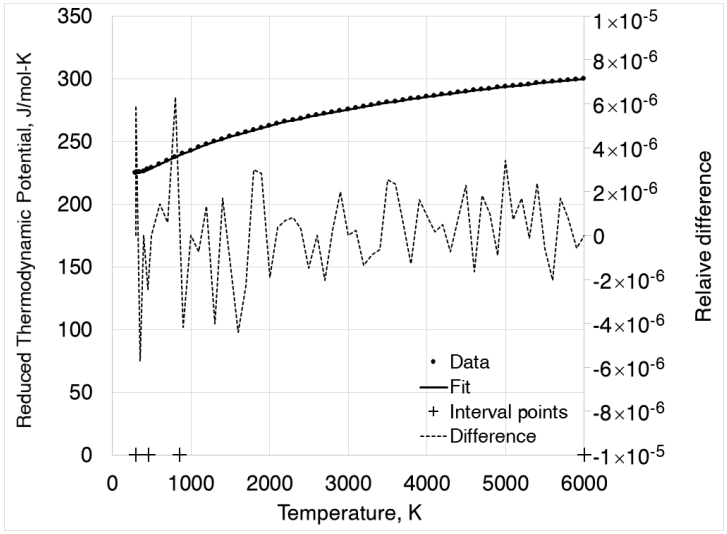
Bromoborane (BBr) – B1Br1(g)

MW = 90.71 g/mol – $\Delta_f H^o$(298.15 K) = 234.304 kJ/mol – S^o(298.15 K) = 224.997 J/mol-K – p^o = 0.1 MPa



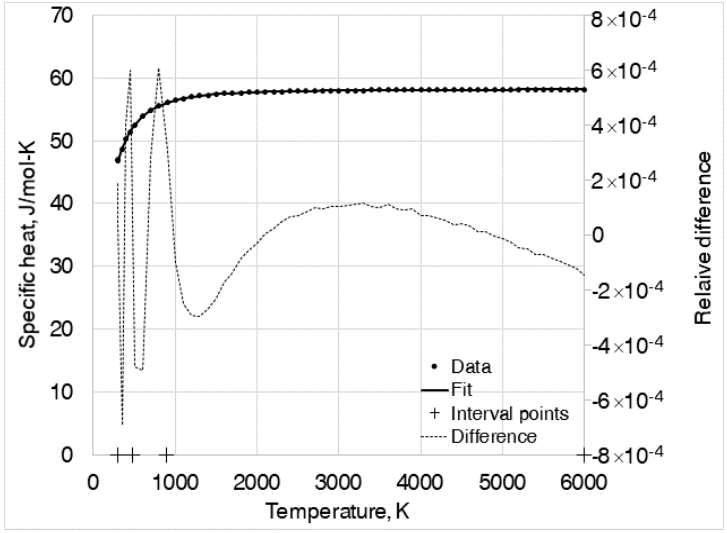


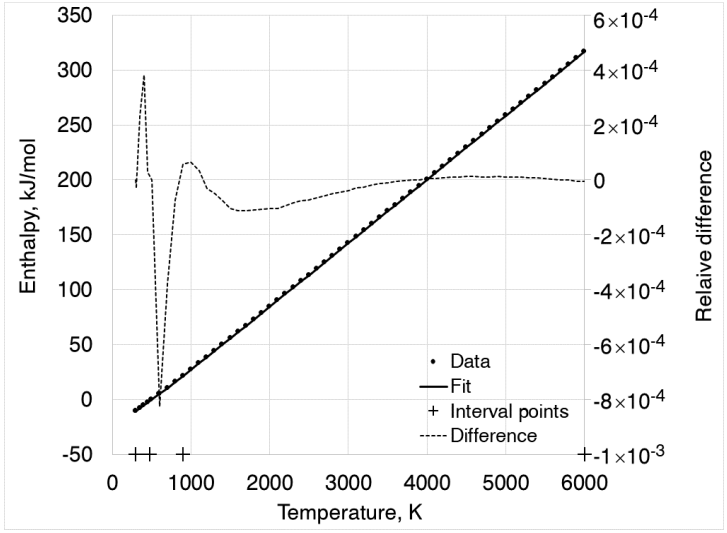


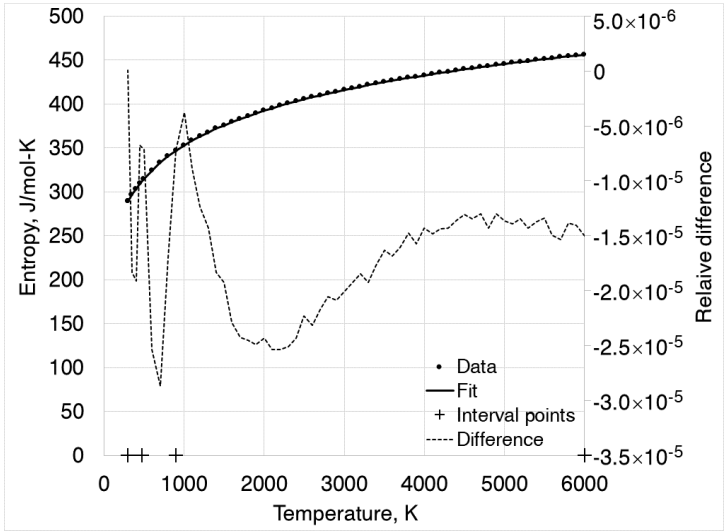


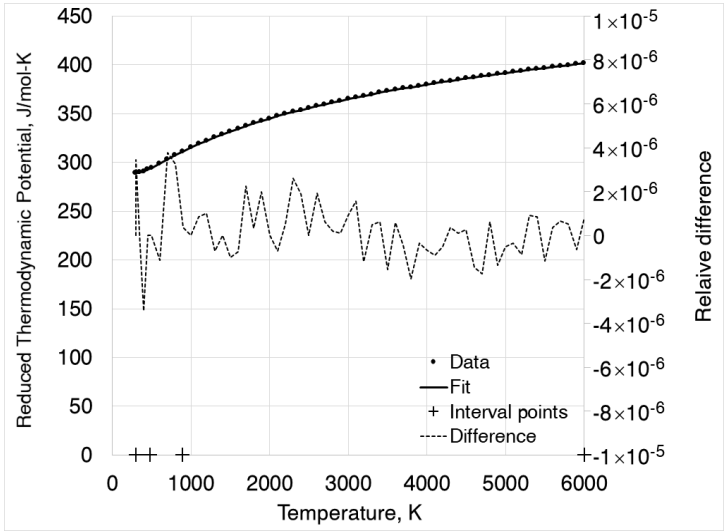
Bromochloroborane (BBrCl) – B1Br1Cl1(g)

MW = 126.17 g/mol – $\Delta_f H^o$(298.15 K) = -10.46 kJ/mol – S^o(298.15 K) = 289.033 J/mol-K – p^o = 0.1 MPa



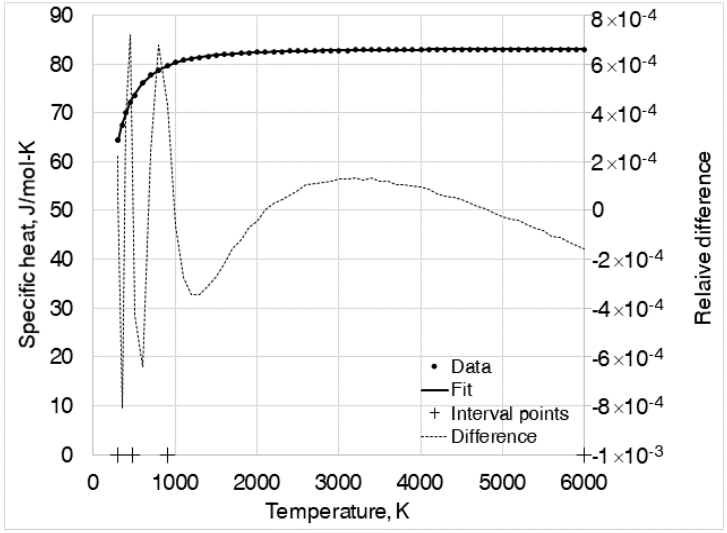


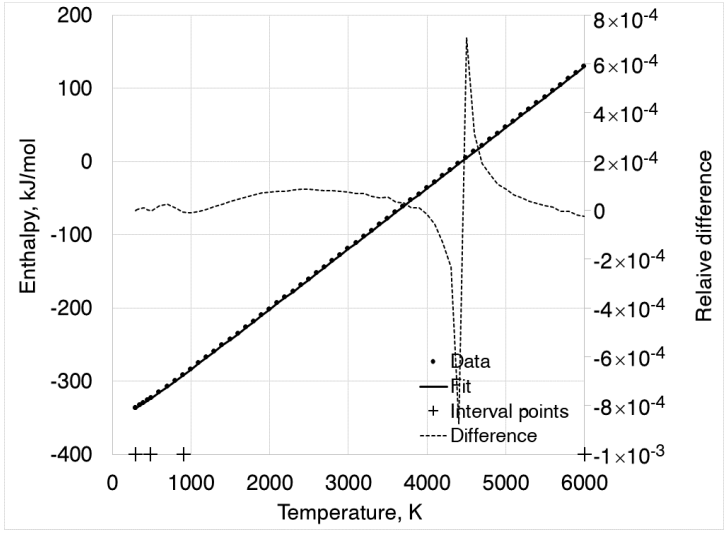


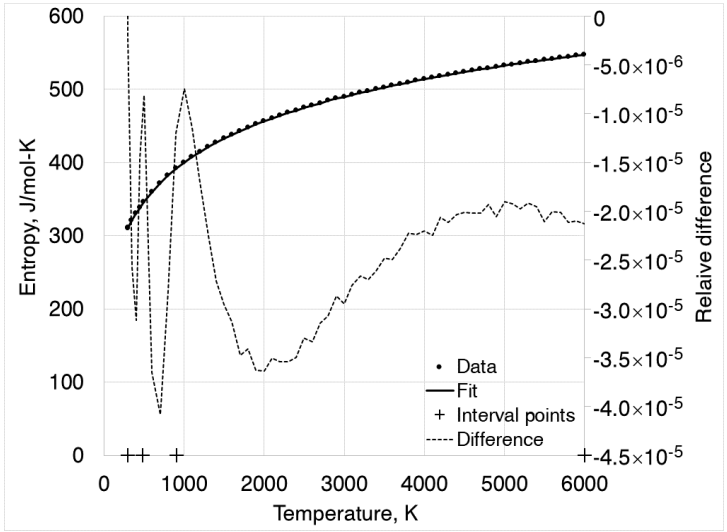


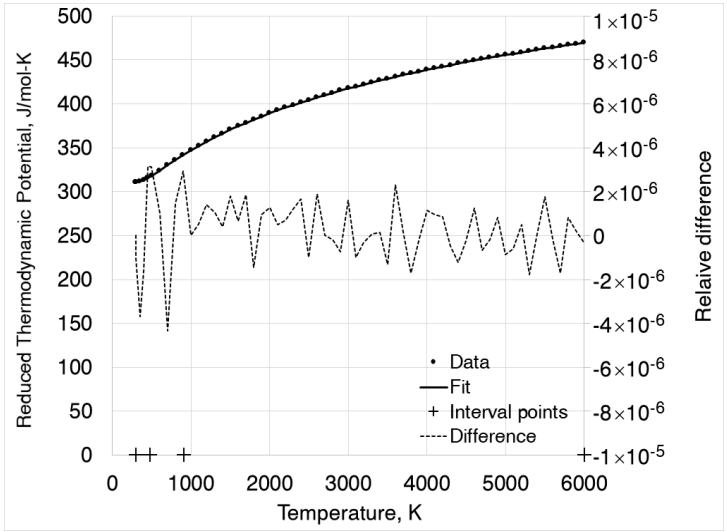
Bromodichloroborane ($BBrCl_2$) – B1Br1Cl2(g)

MW = 161.62 g/mol – $\Delta_f H^o$(298.15 K) = -336.812 kJ/mol – S^o(298.15 K) = 310.4 J/mol-K – p^o = 0.1 MPa



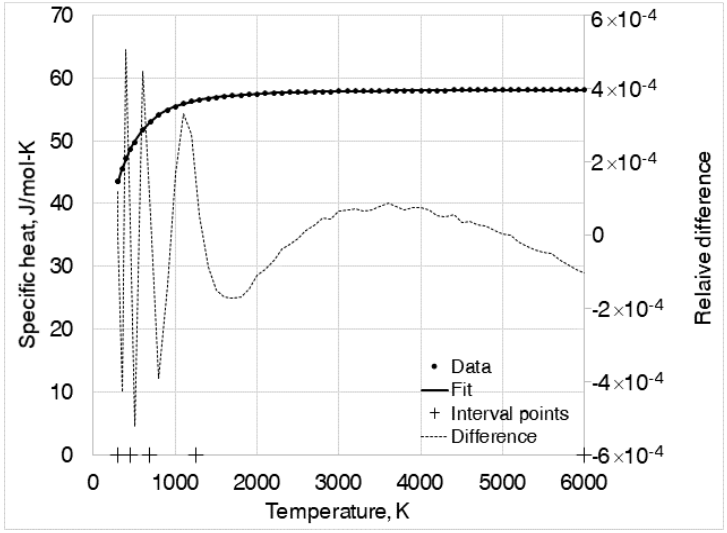


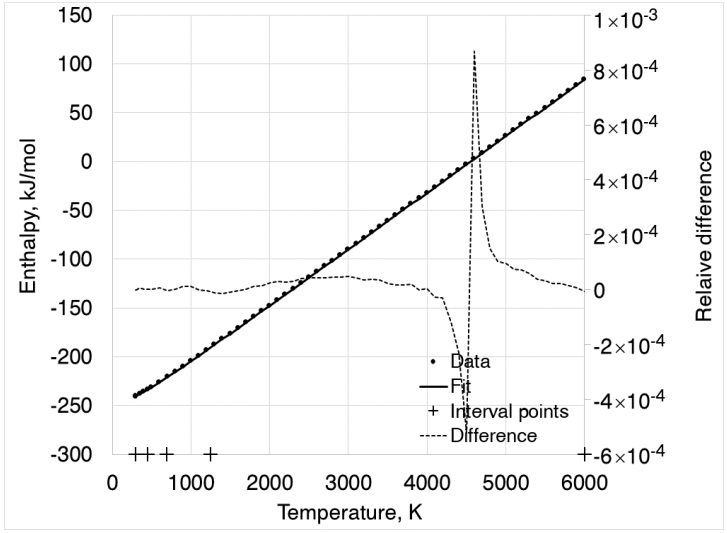


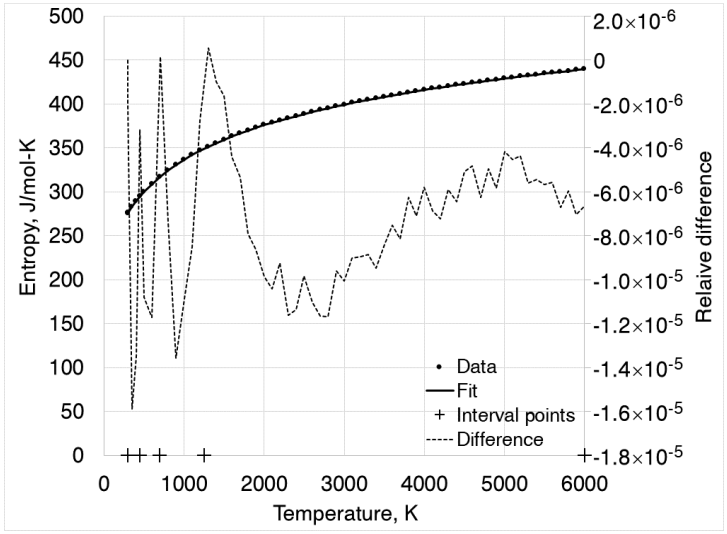


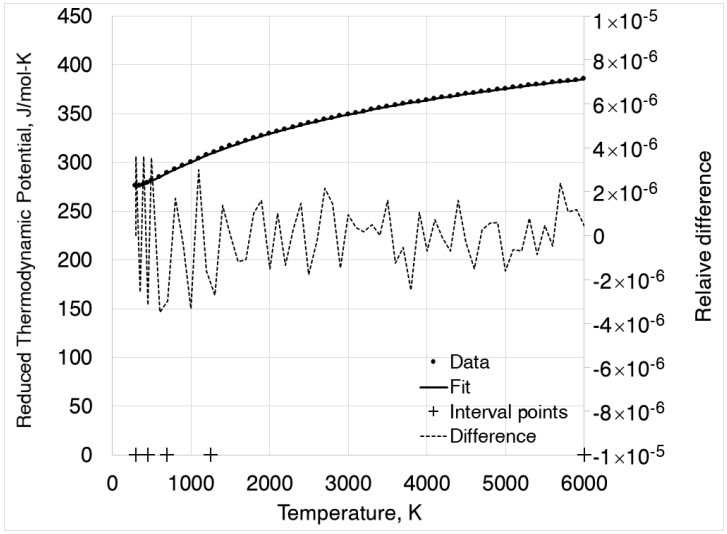
Bromofluoroborane (BBrF) – B1Br1F1(g)

MW = 109.71 g/mol – $\Delta_f H^o$(298.15 K) = -240.58 kJ/mol – S^o(298.15 K) = 275.77 J/mol-K – p^o = 0.1 MPa



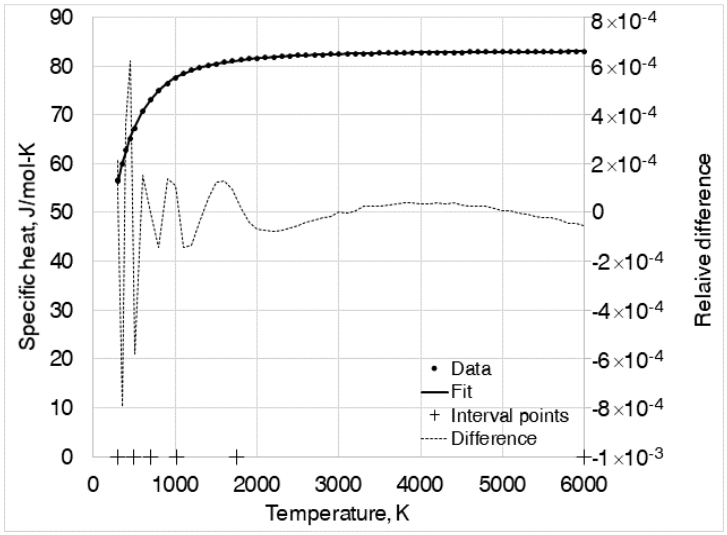


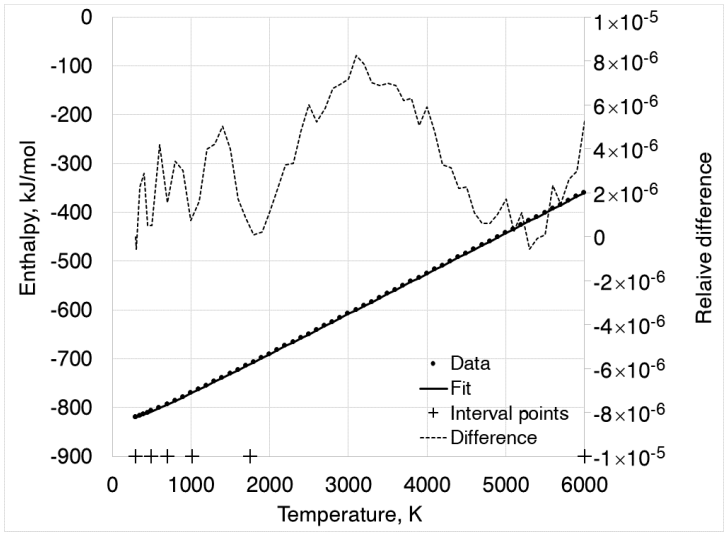


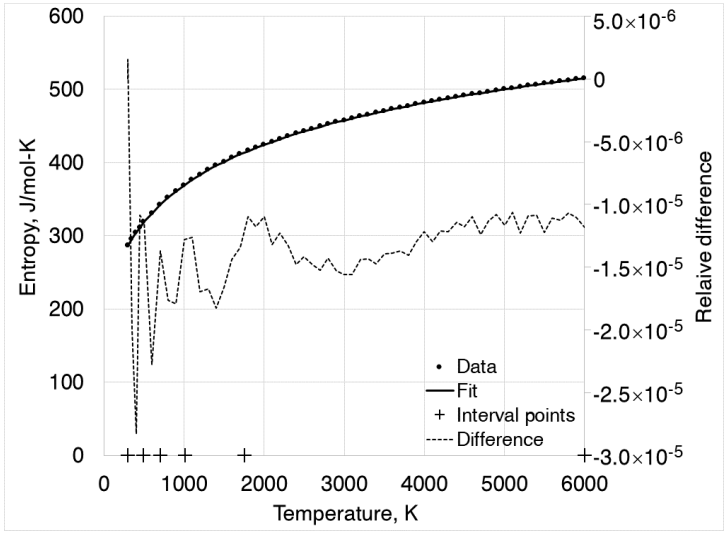


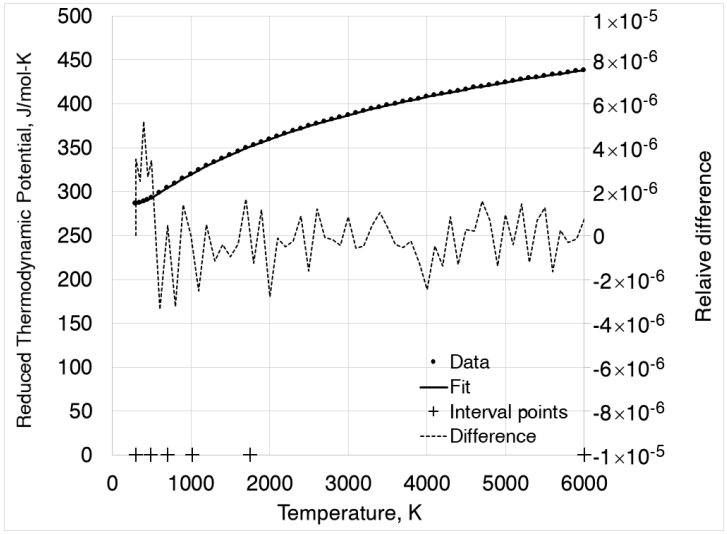
Bromodifluoroborane ($BBrF_2$) – B1Br1F2(g)

MW = 128.71 g/mol – $\Delta_f H^o$(298.15 K) = -820.064 kJ/mol – S^o(298.15 K) = 286.387 J/mol-K – p^o = 0.1 MPa



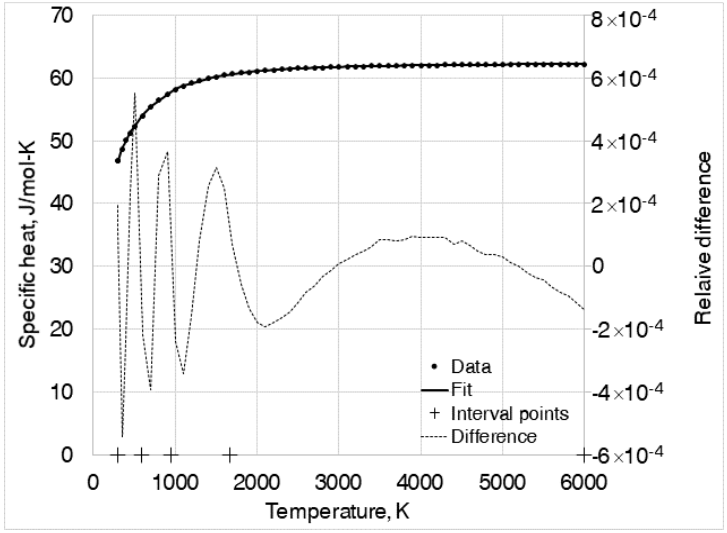


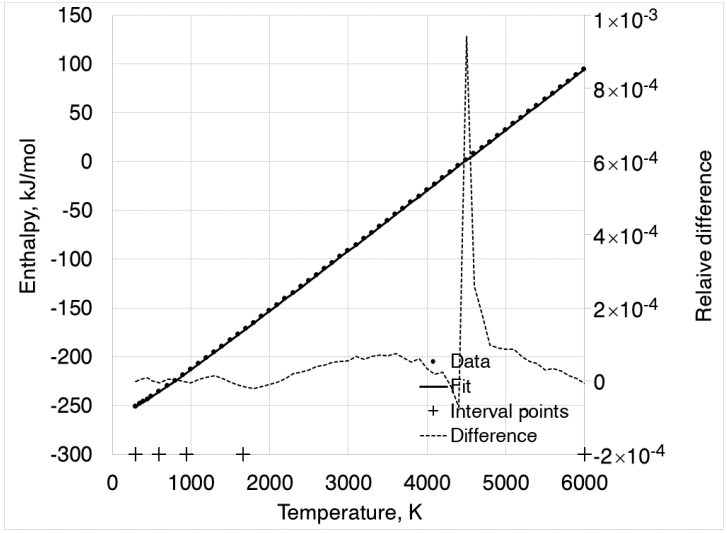


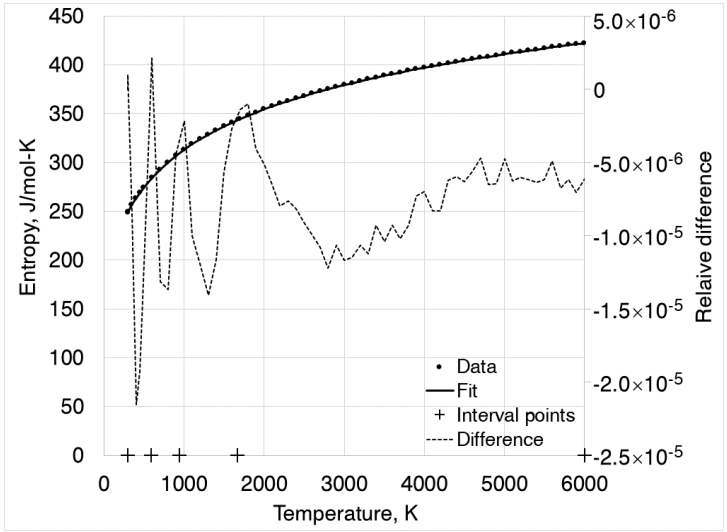


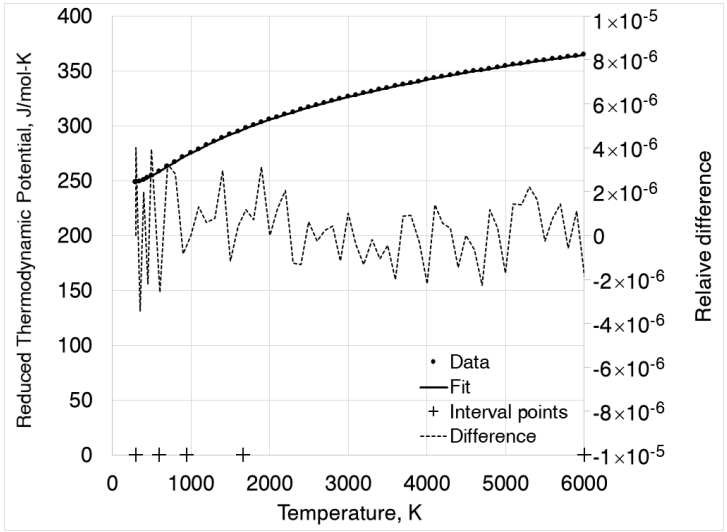
Boron Bromide Oxide (OBBr) – B1Br1O1(g)

MW = 106.71 g/mol – $\Delta_f H^o$(298.15 K) = -251.04 kJ/mol – S^o(298.15 K) = 248.895 J/mol-K – p^o = 0.1 MPa



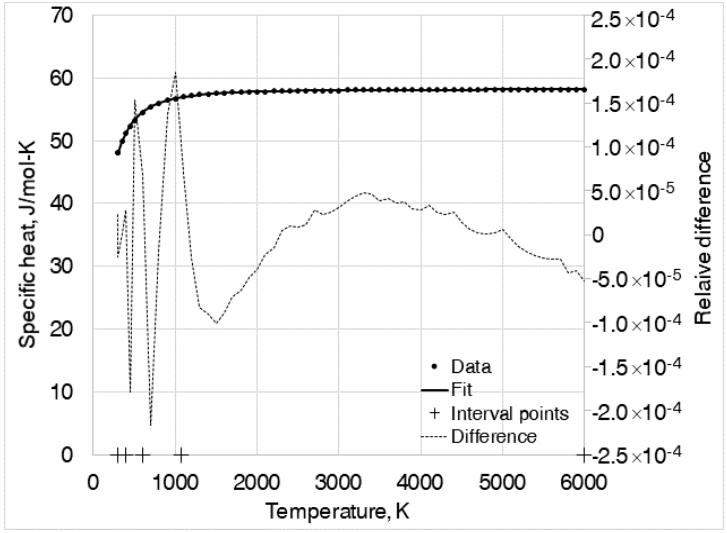


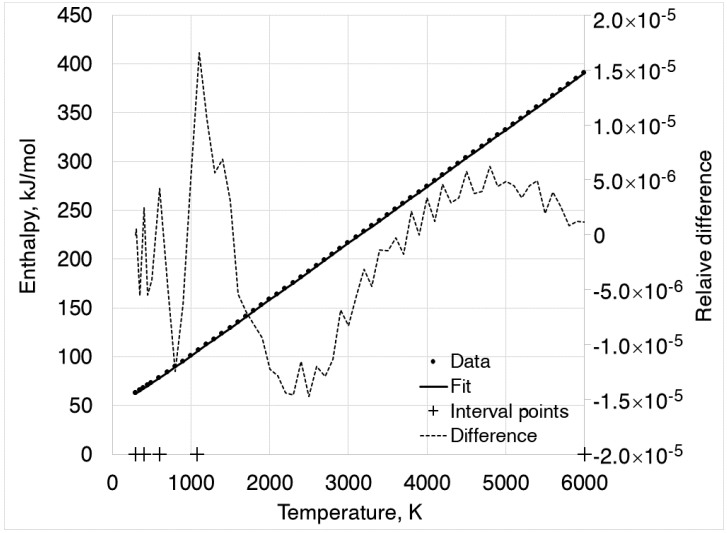


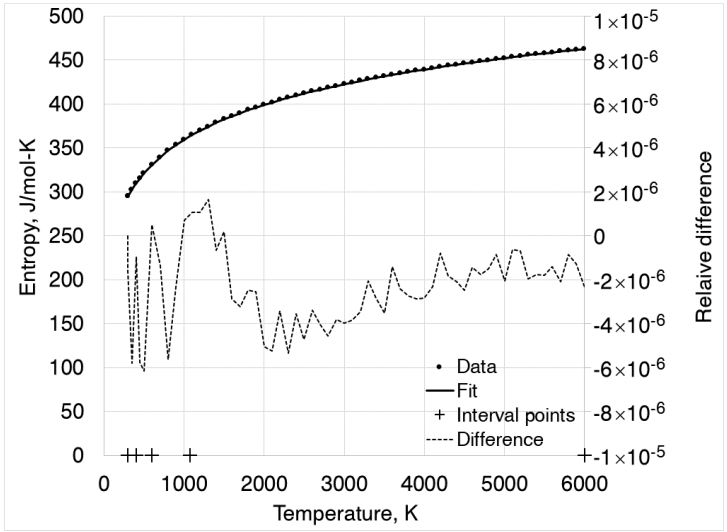


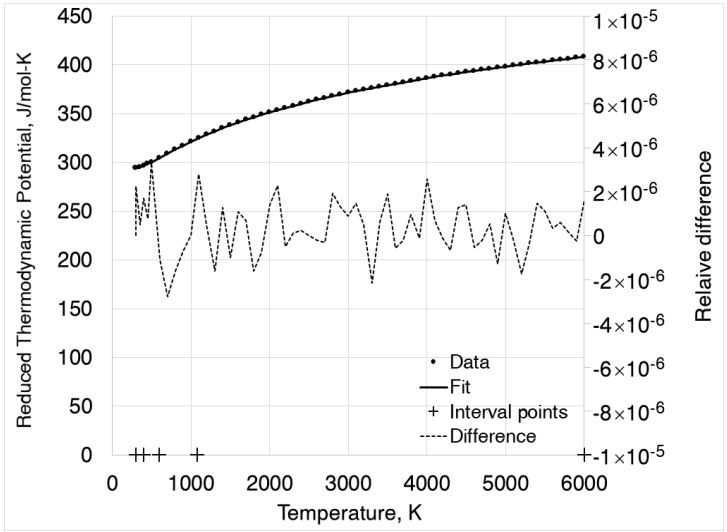
Dibromoborane (BBr_2) – B1Br2(g)

MW = 170.62 g/mol – $\Delta_f H^o$(298.15 K) = 62.76 kJ/mol – S^o(298.15 K) = 294.662 J/mol-K – p^o = 0.1 MPa



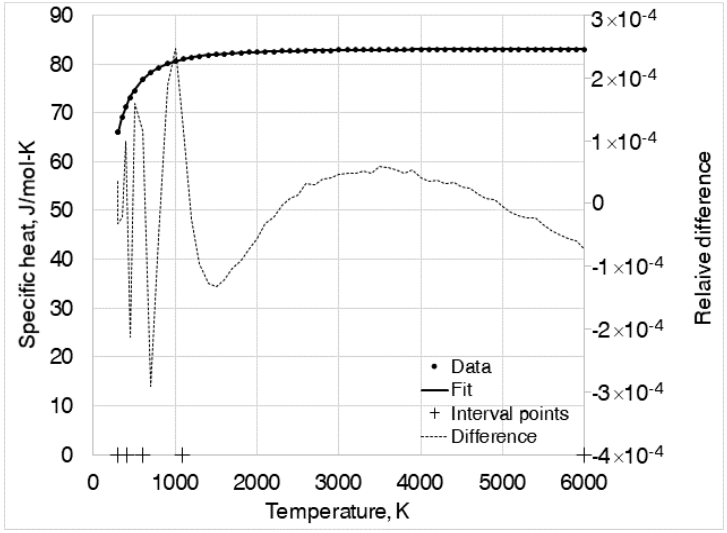


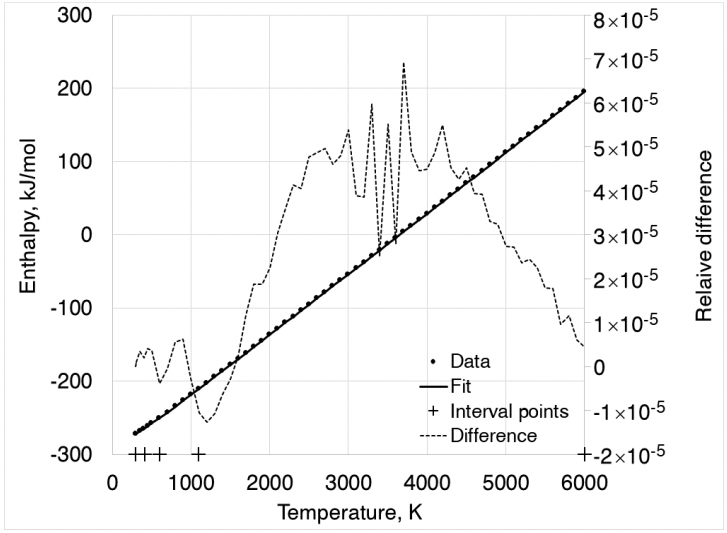


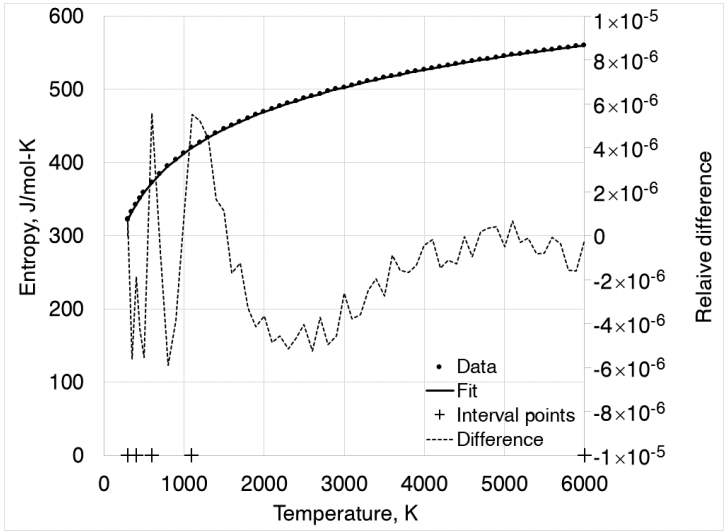


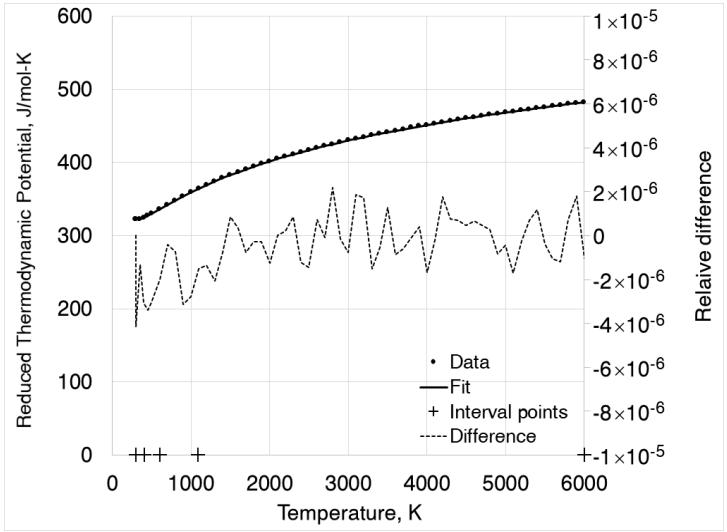
Dibromochloroborane (BBr_2Cl) – B1Br2Cl1(g)

MW = 206.07 g/mol – $\Delta_f H^o$(298.15 K) = -271.96 kJ/mol – S^o(298.15 K) = 321.889 J/mol-K – p^o = 0.1 MPa



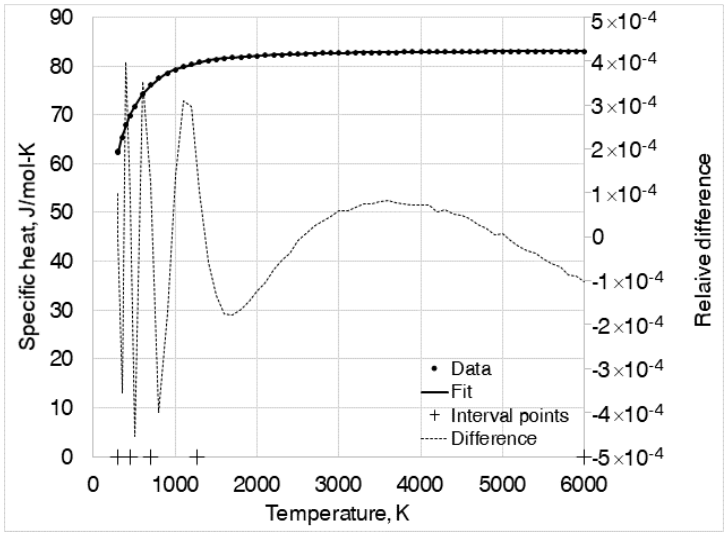


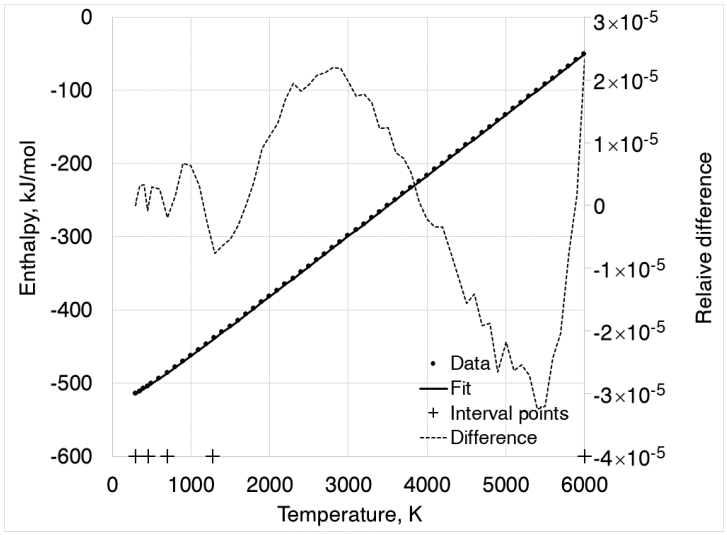


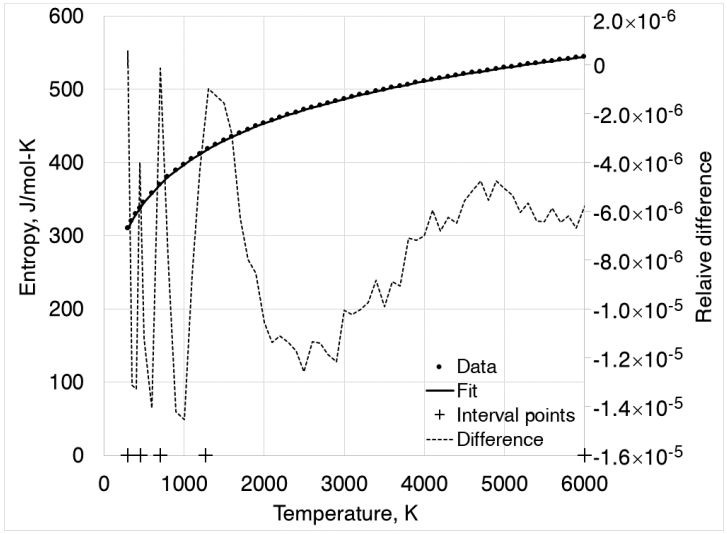


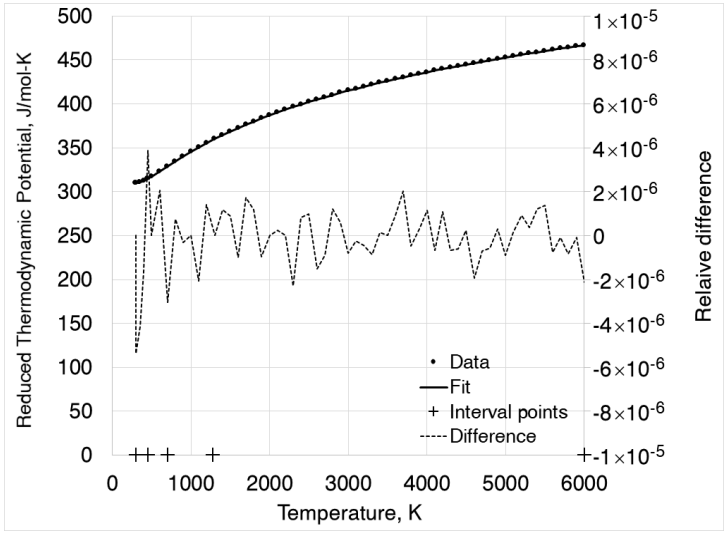
Dibromofluoroborane (BBr_2F) – B1Br2F1(g)

MW = 189.62 g/mol – $\Delta_f H^o$(298.15 K) = -514.632 kJ/mol – S^o(298.15 K) = 309.994 J/mol-K – p^o = 0.1 MPa



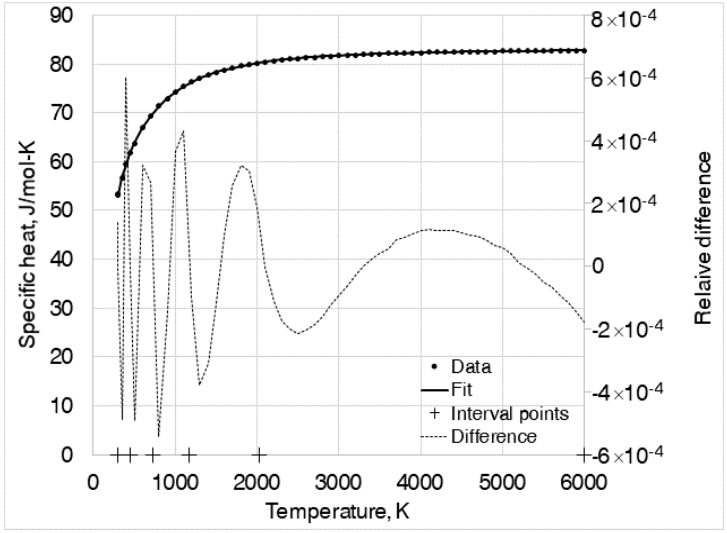


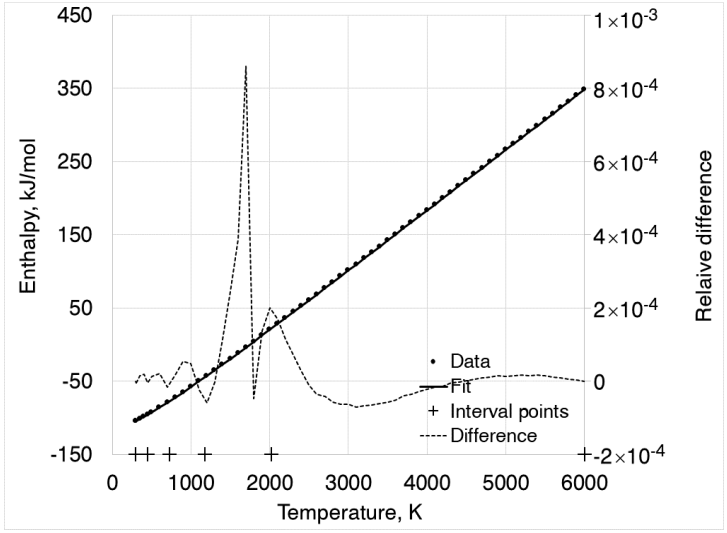


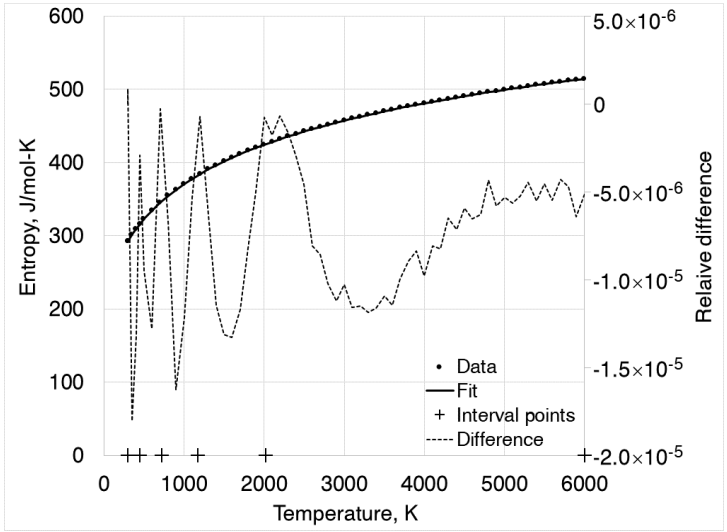


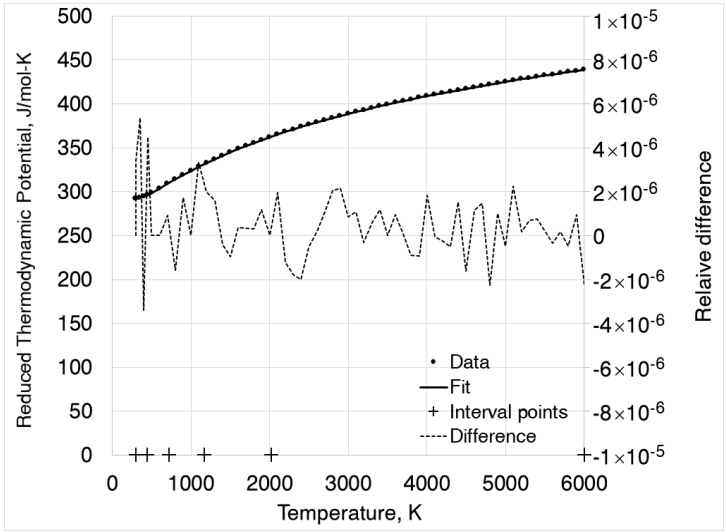
Dibromoborane ($BHBr_2$) – B1Br2H1(g)

MW = 171.63 g/mol – $\Delta_f H^o$(298.15 K) = -104.6 kJ/mol – S^o(298.15 K) = 292.265 J/mol-K – p^o = 0.1 MPa



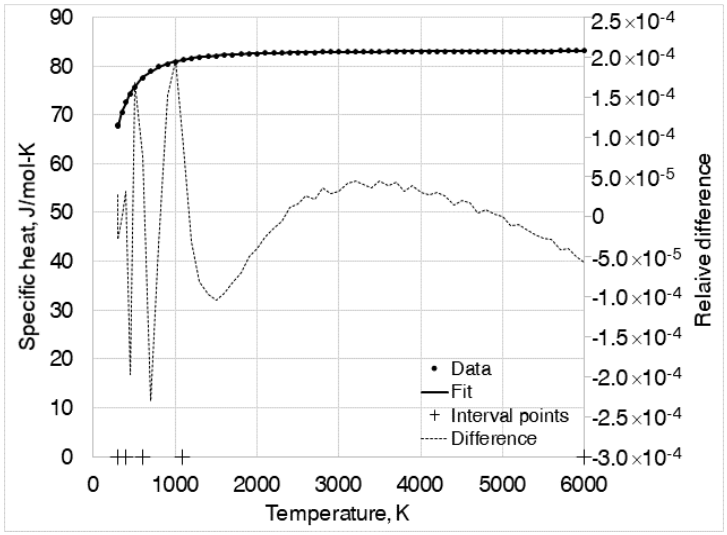


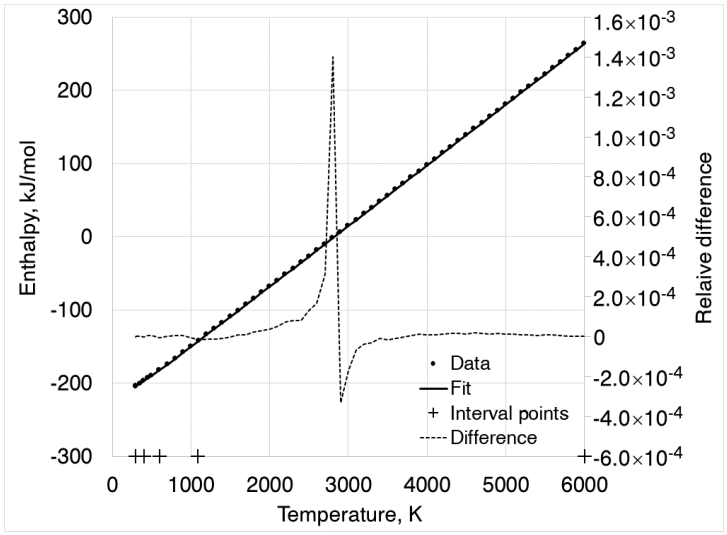


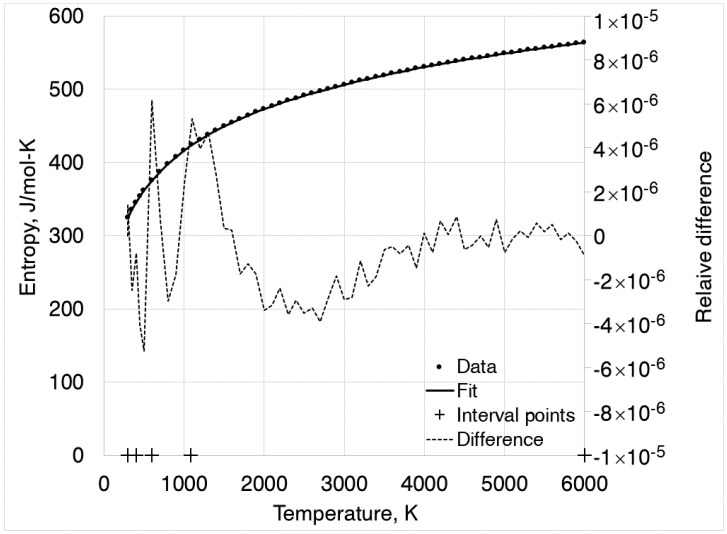


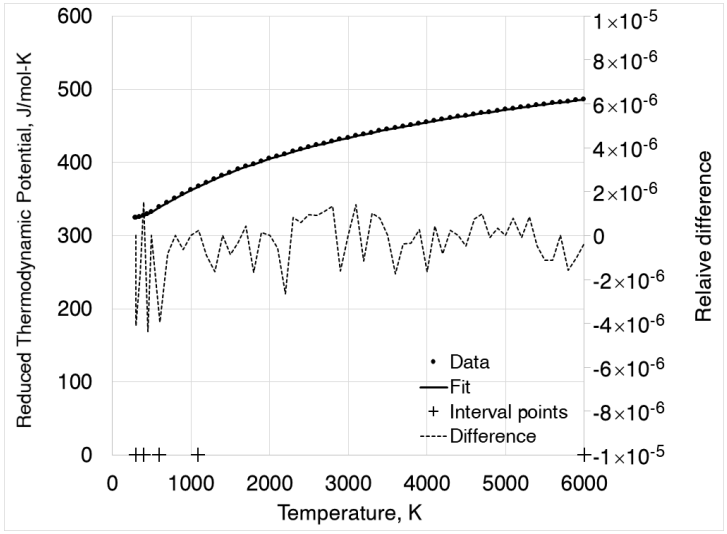
Tribromoborane (BBr_3) – B1Br3(g)

MW = 250.52 g/mol – $\Delta_f H^o$(298.15 K) = -204.179 kJ/mol – S^o(298.15 K) = 324.307 J/mol-K – p^o = 0.1 MPa



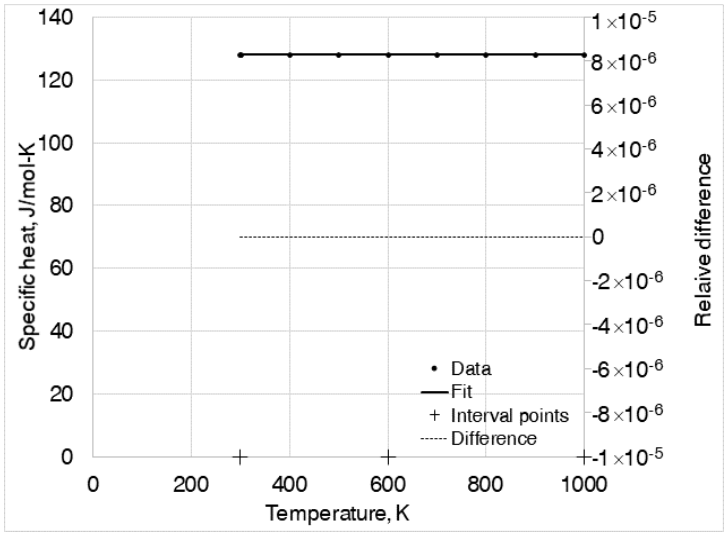


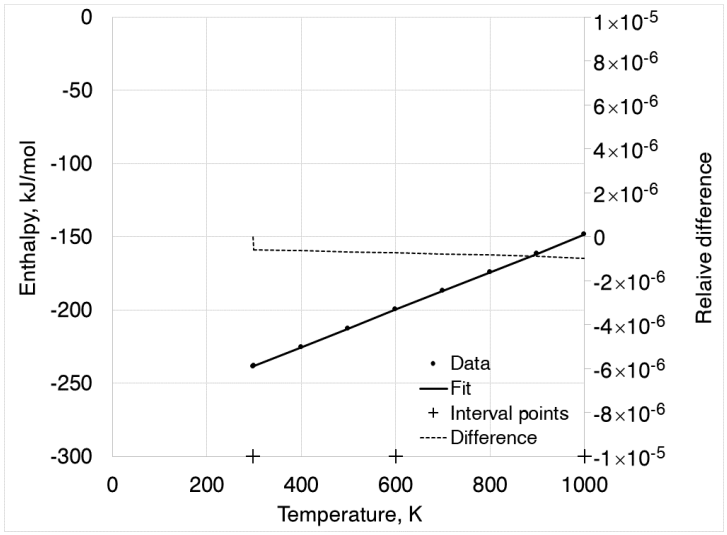


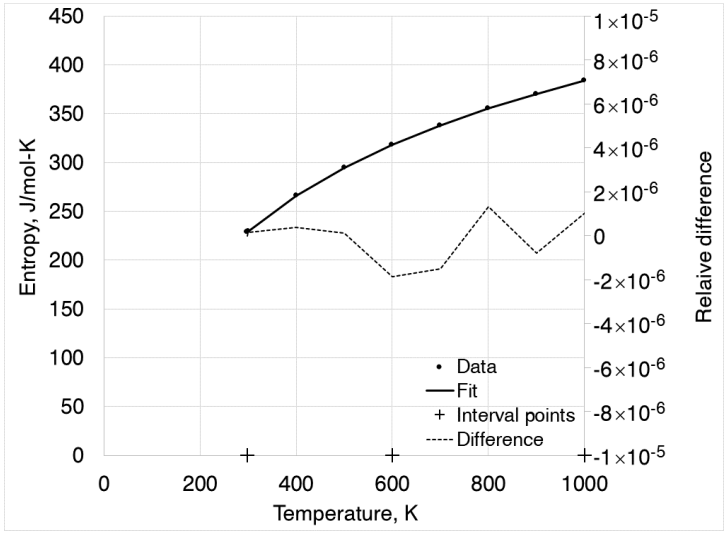


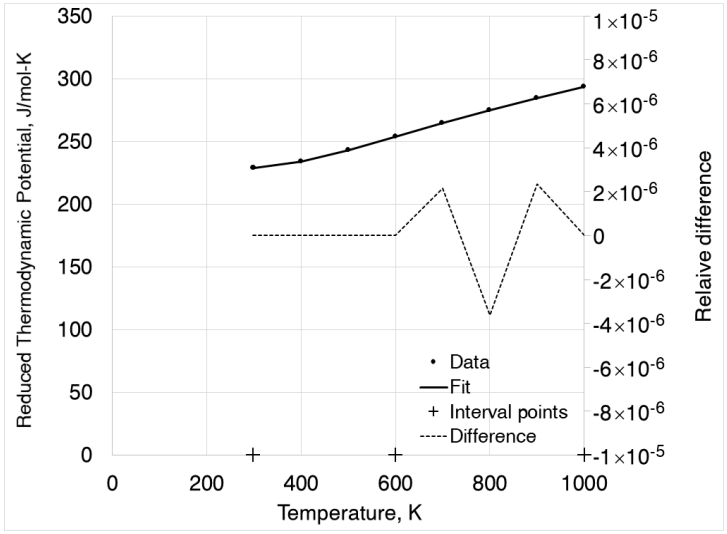
Tribromoborane (BBr_3) – B1Br3(l)

MW = 250.52 g/mol – $\Delta_f H^o$(298.15 K) = -238.488 kJ/mol – S^o(298.15 K) = 228.865 J/mol-K – p^o = 0.1 MPa



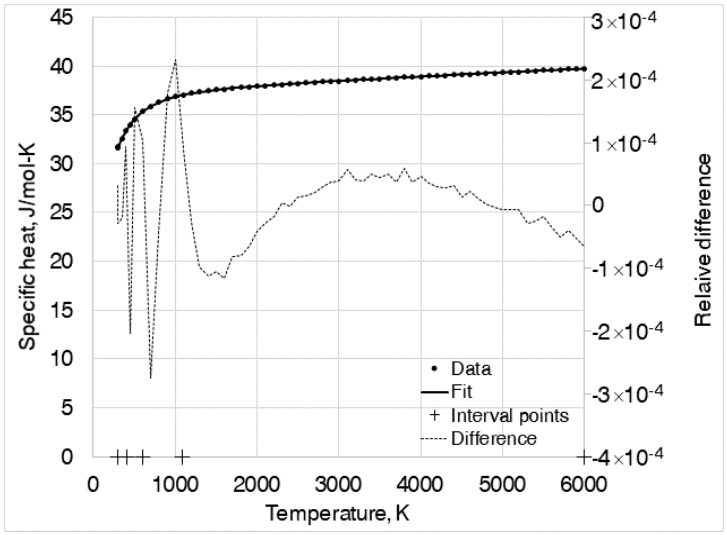


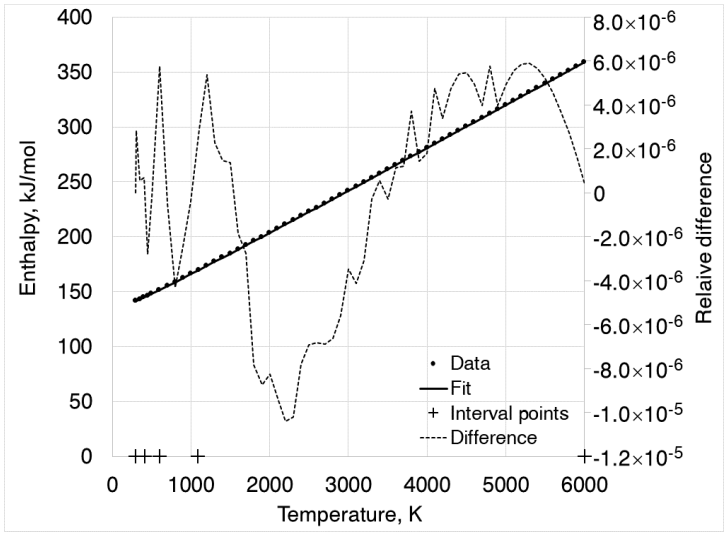


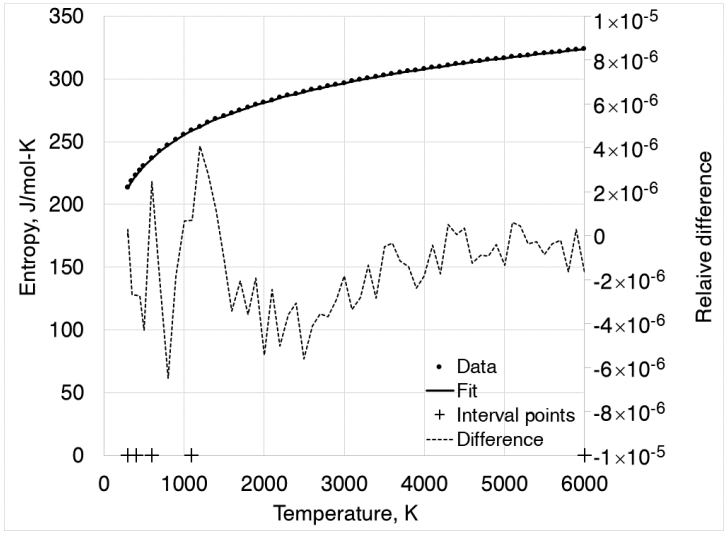


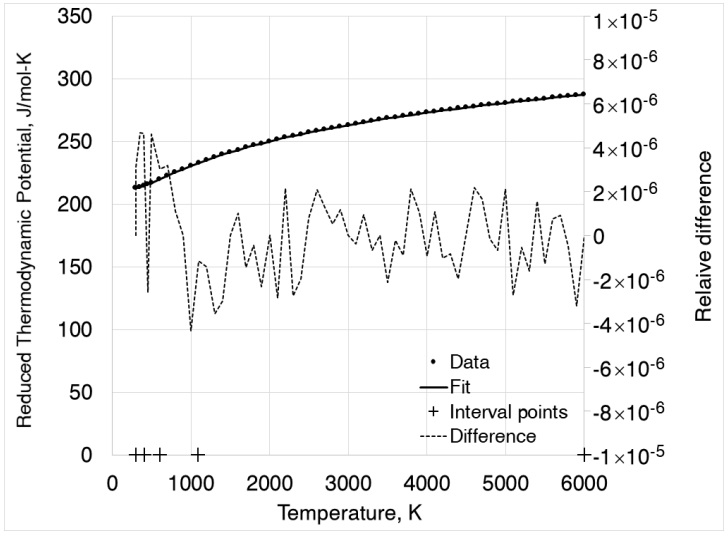
Chloroborane (BCl) – B1Cl1(g)

MW = 46.26 g/mol – $\Delta_f H^o$(298.15 K) = 141.419 kJ/mol – S^o(298.15 K) = 213.245 J/mol-K – p^o = 0.1 MPa



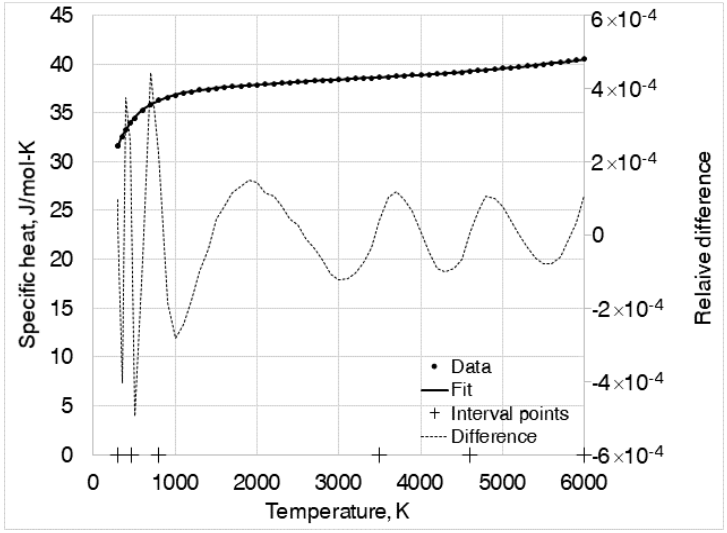


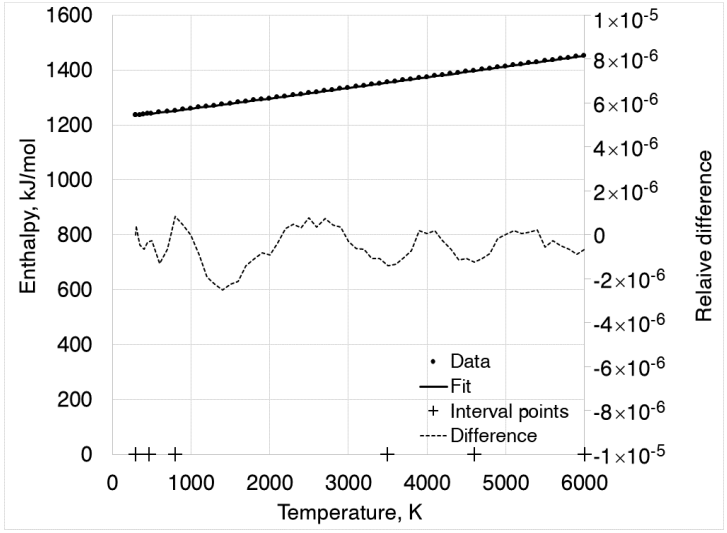


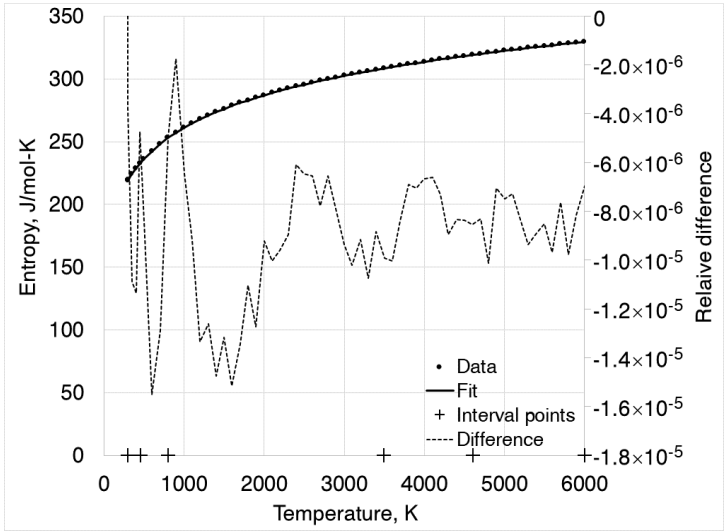


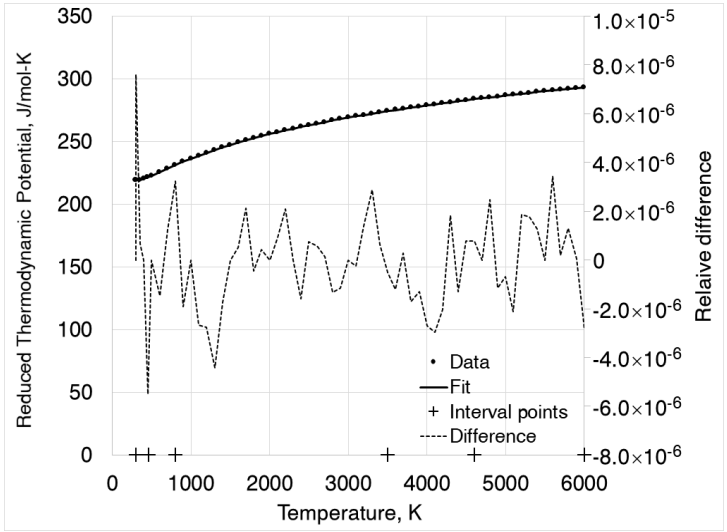
Chloroborane, Ion (BCl^+) – B1Cl1+(g)

MW = 46.26 g/mol – $\Delta_f H^o$(298.15 K) = 1234.28 kJ/mol – S^o(298.15 K) = 219.131 J/mol-K – p^o = 0.1 MPa



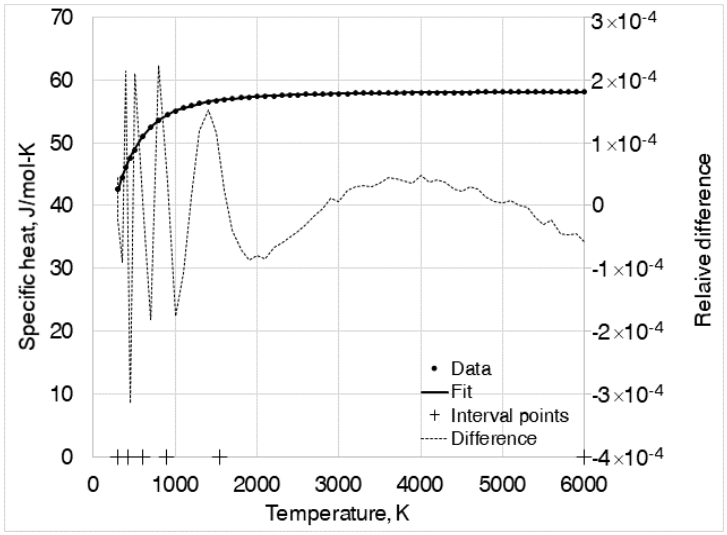


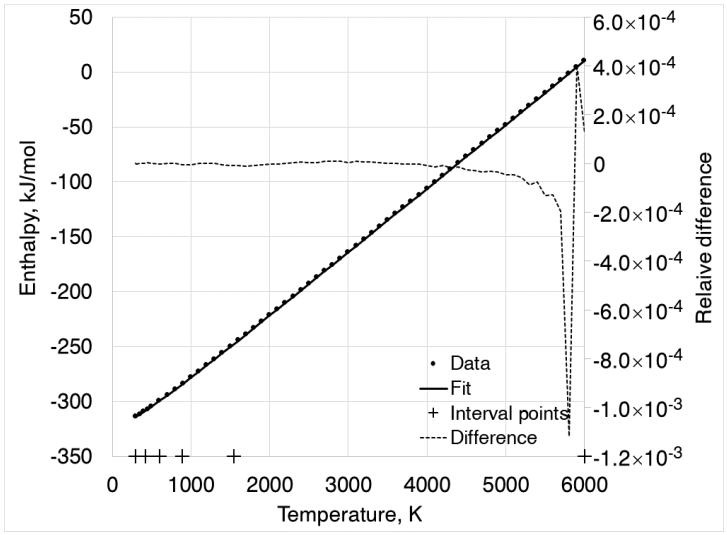


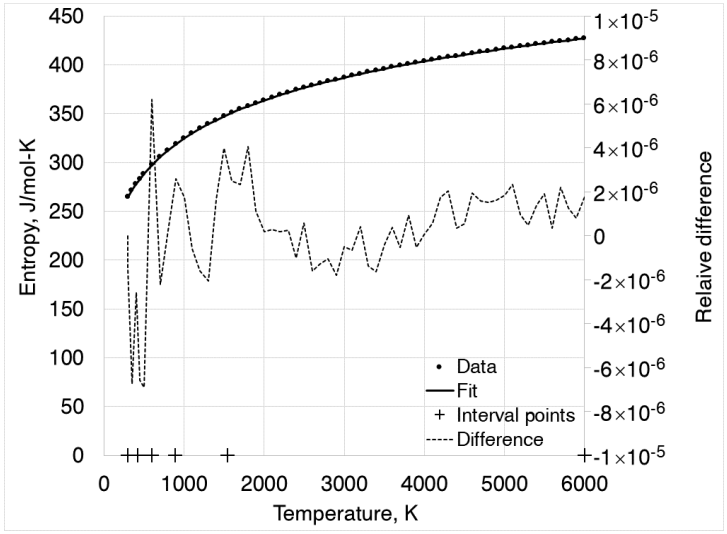


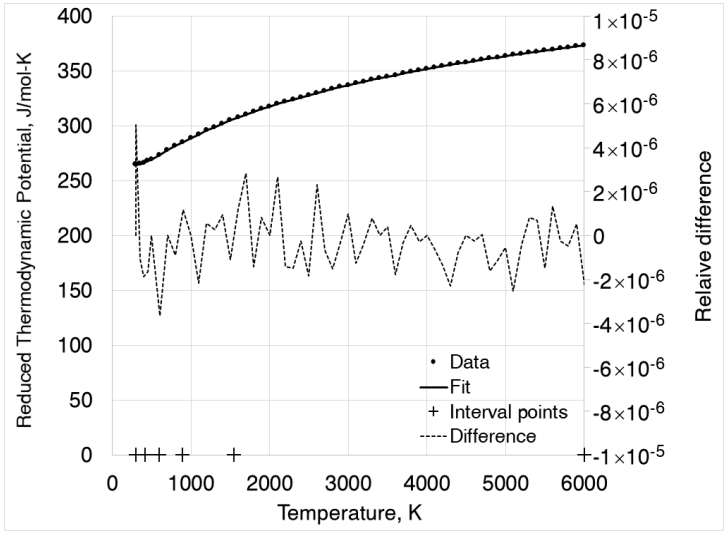
Chlorofluoroborane (BClF) – B1Cl1F1(g)

MW = 65.26 g/mol – $\Delta_f H^o$(298.15 K) = -313.8 kJ/mol – S^o(298.15 K) = 264.65 J/mol-K – p^o = 0.1 MPa



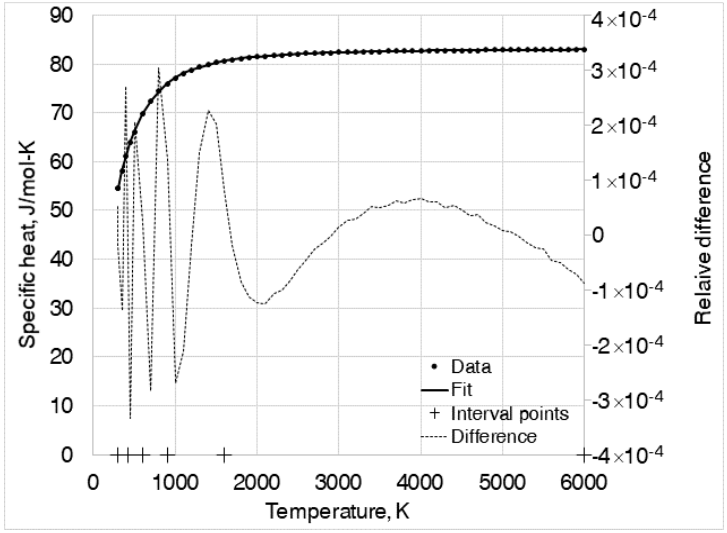


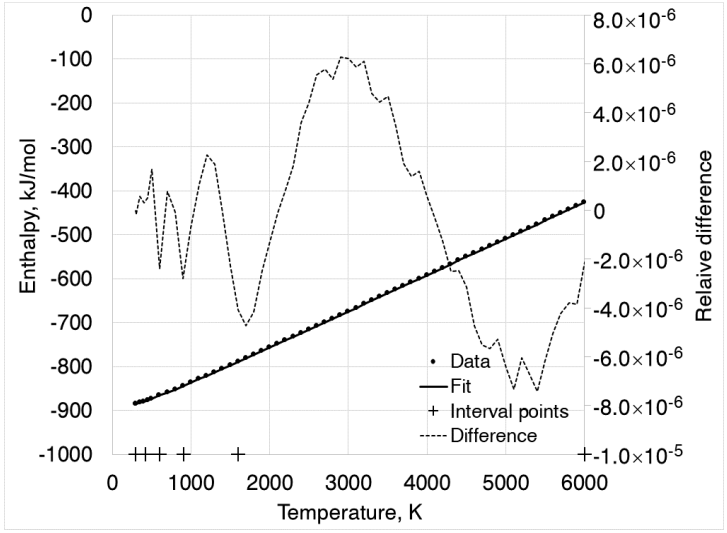


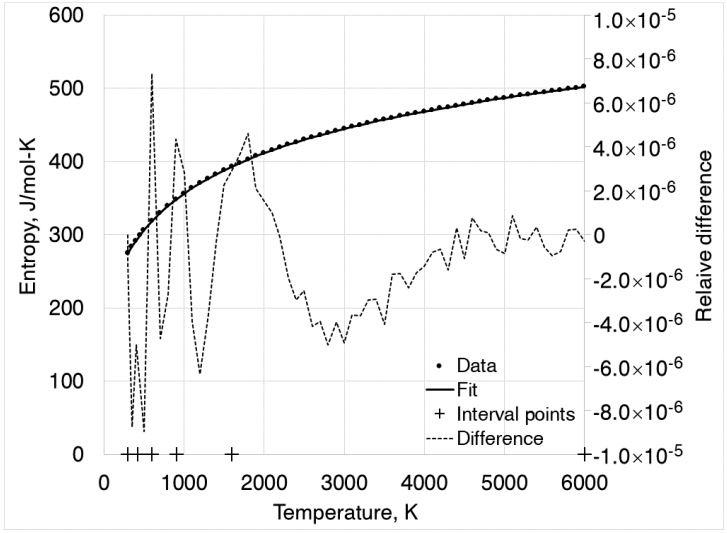


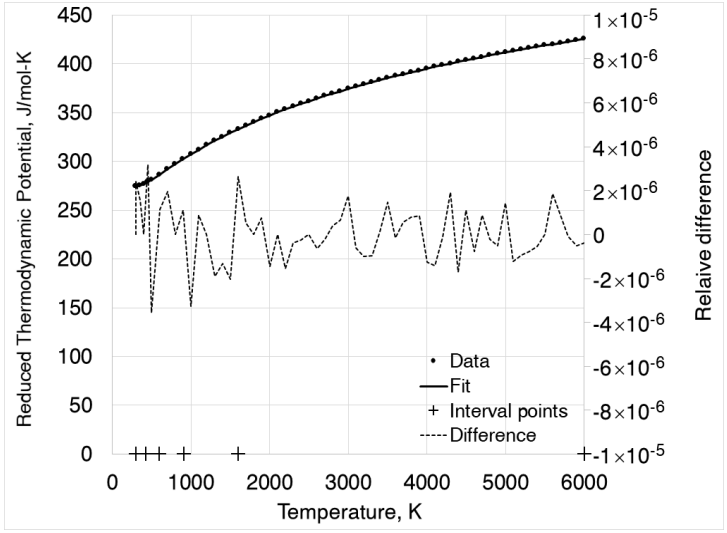
Chlorodifluoroborane ($BClF_2$) – B1Cl1F2(g)

MW = 84.26 g/mol – $\Delta_f H^o$(298.15 K) = -885.334 kJ/mol – S^o(298.15 K) = 274.83 J/mol-K – p^o = 0.1 MPa



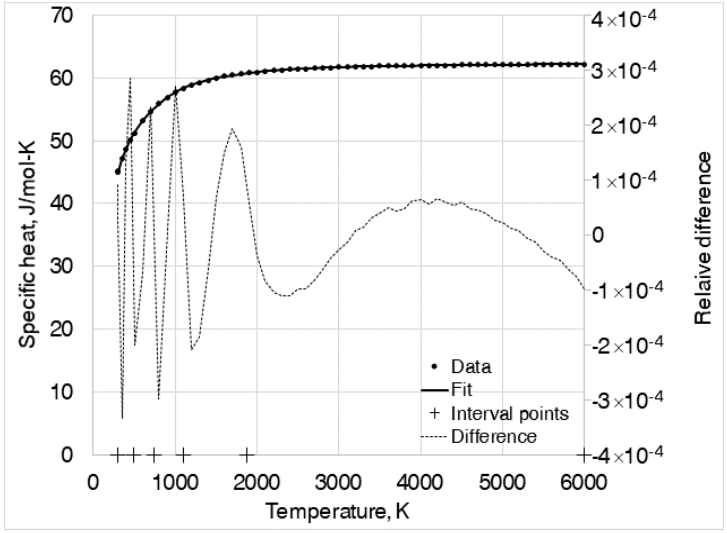


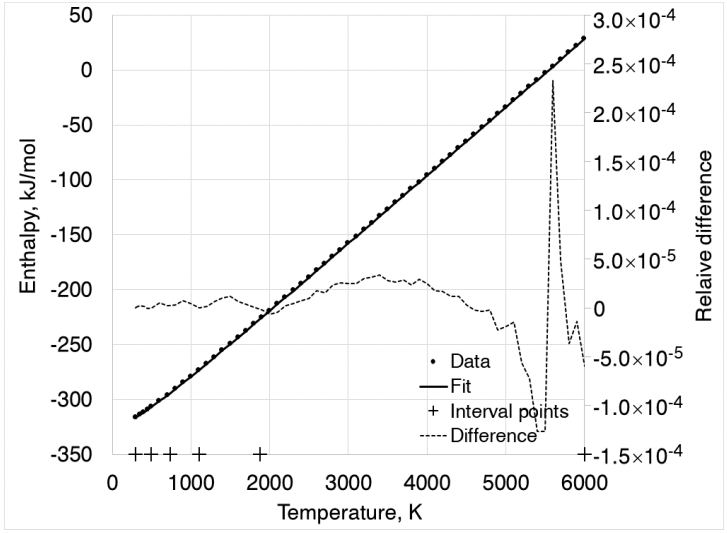


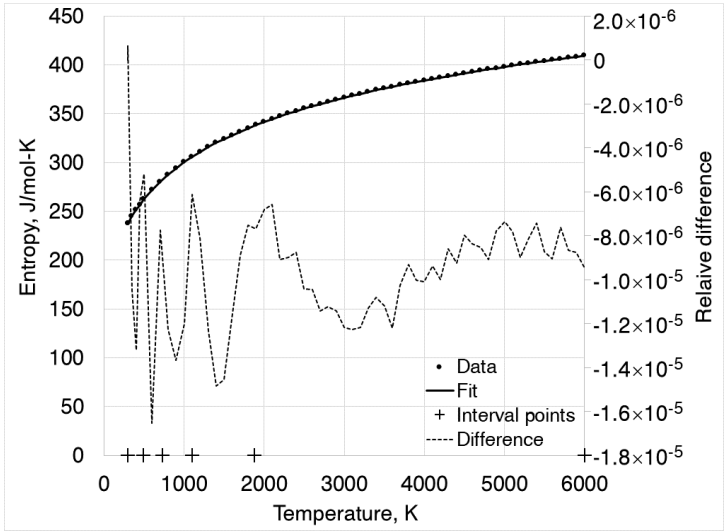


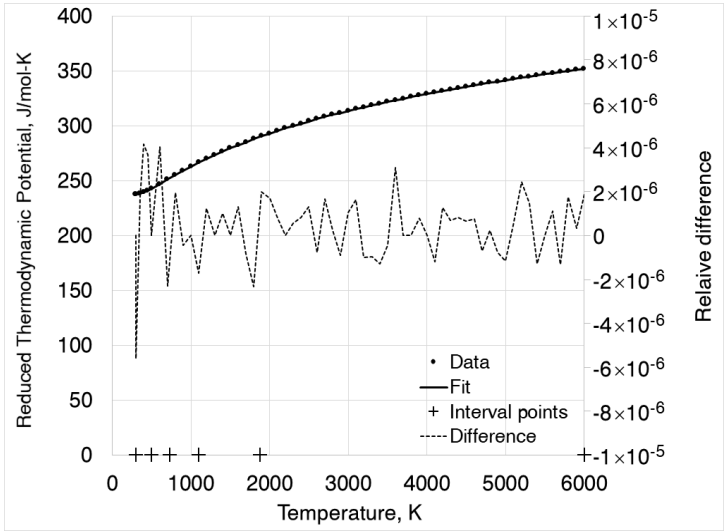
Boron Chloride Oxide (OBCl) – B1Cl1O1(g)

MW = 62.26 g/mol – $\Delta_f H^o$(298.15 K) = -316.31 kJ/mol – S^o(298.15 K) = 237.417 J/mol-K – p^o = 0.1 MPa



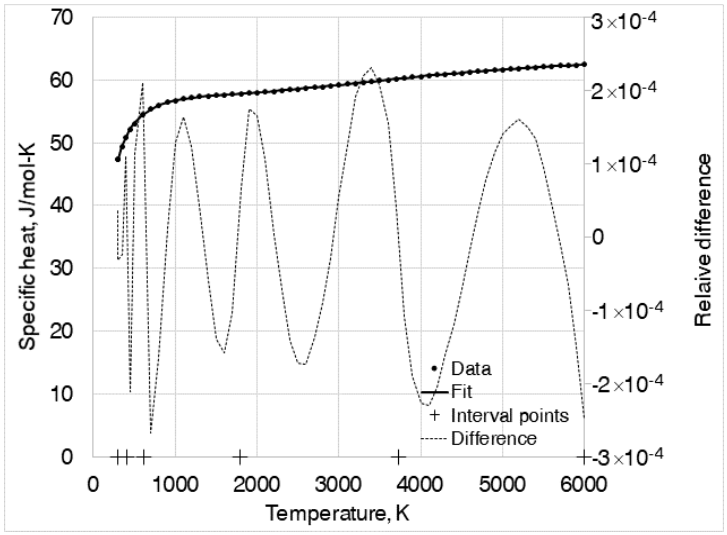


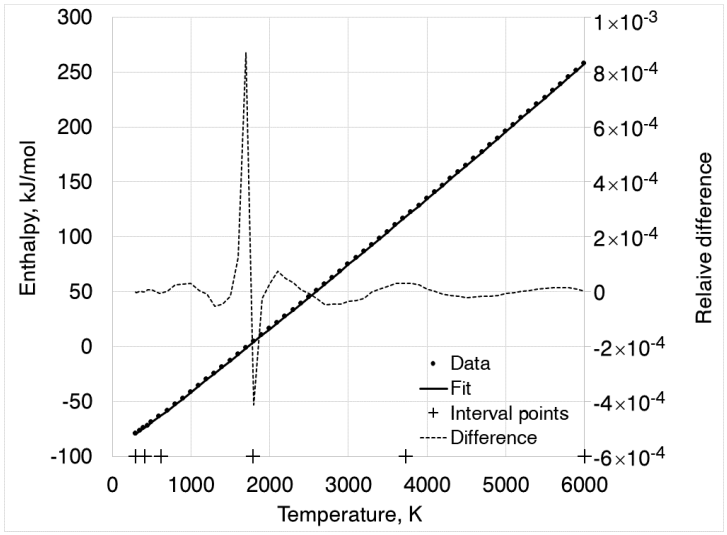


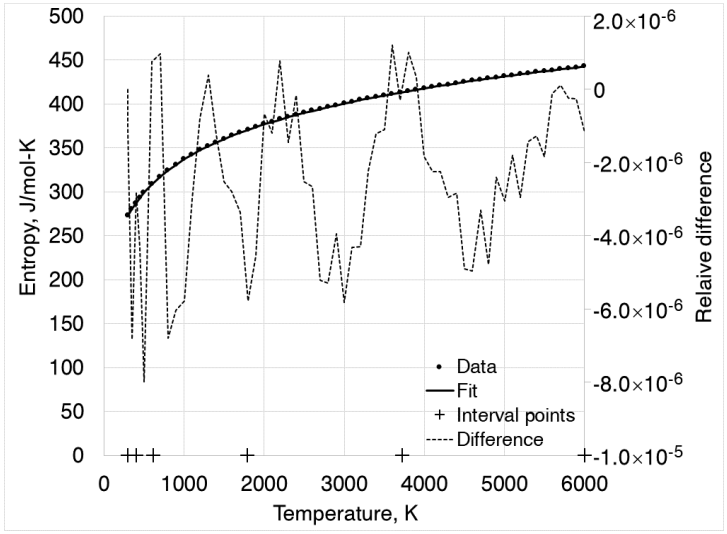


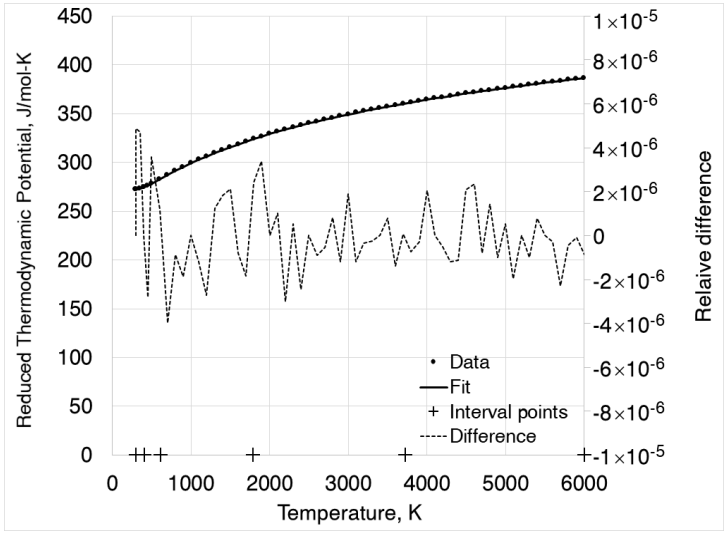
Dichloroborane (BCl_2) – B1Cl2(g)

MW = 81.72 g/mol – $\Delta_f H^o$(298.15 K) = -79.496 kJ/mol – S^o(298.15 K) = 272.677 J/mol-K – p^o = 0.1 MPa



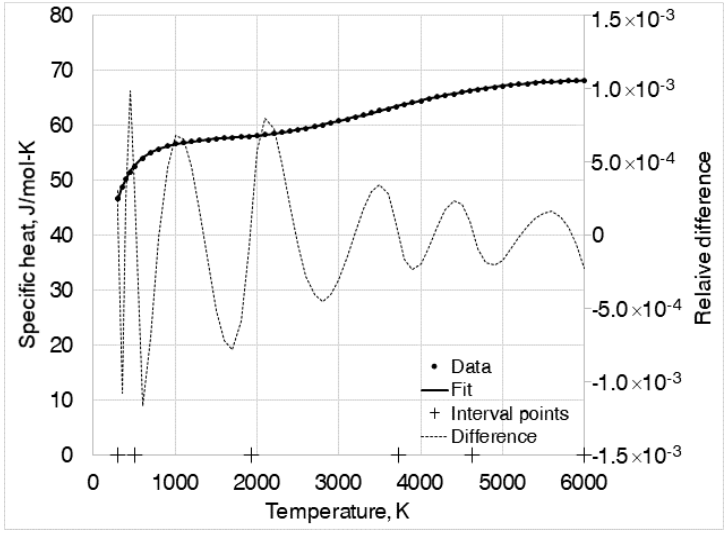


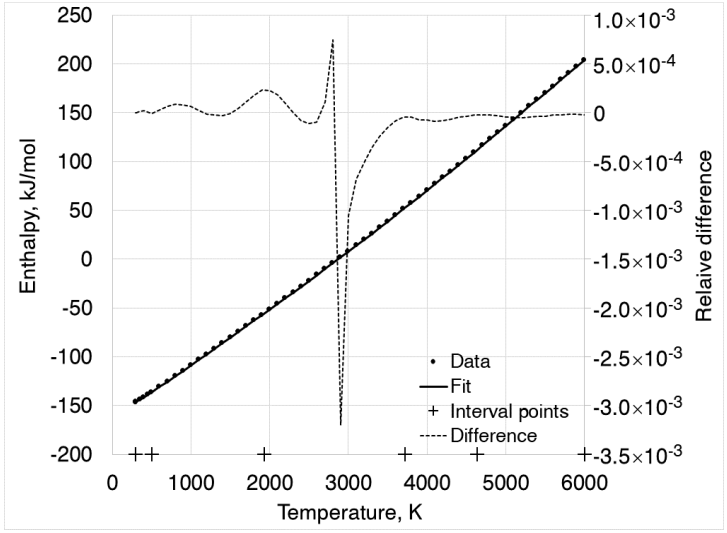


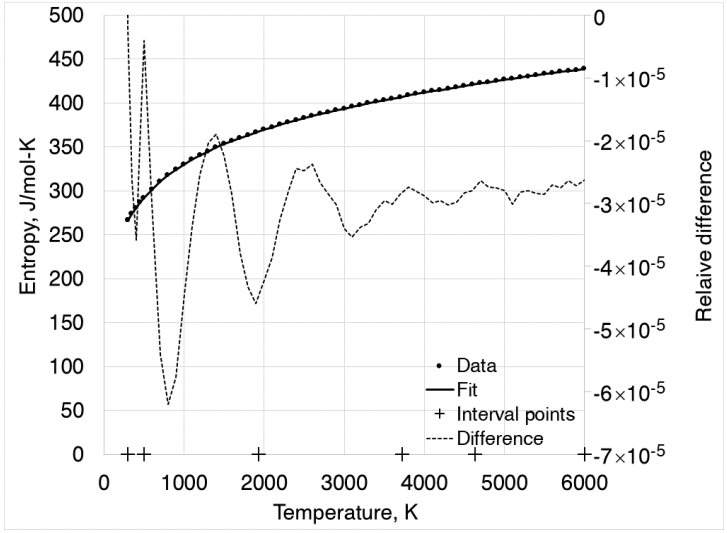


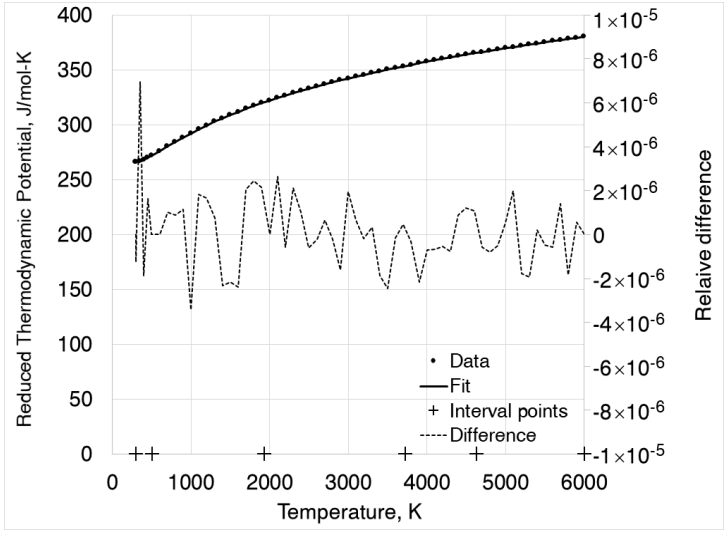
Dichloroborane, Ion (BCl_2^-) – B1Cl2-(g)

MW = 81.72 g/mol – $\Delta_f H^o$(298.15 K) = -146.44 kJ/mol – S^o(298.15 K) = 266.251 J/mol-K – p^o = 0.1 MPa



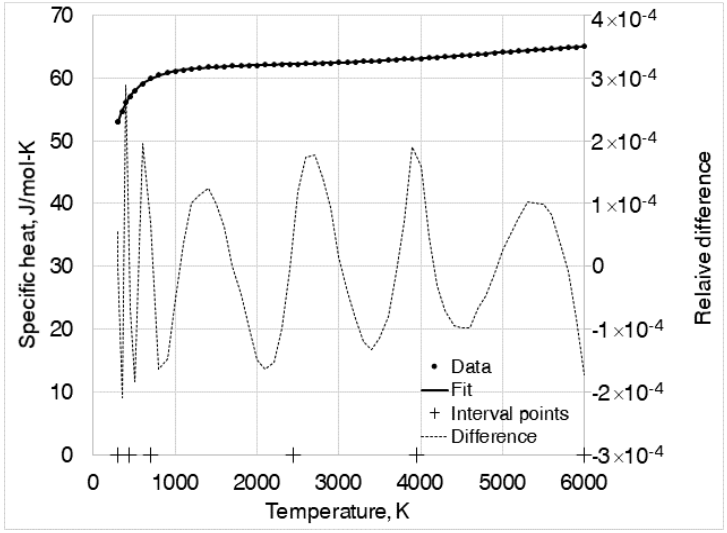


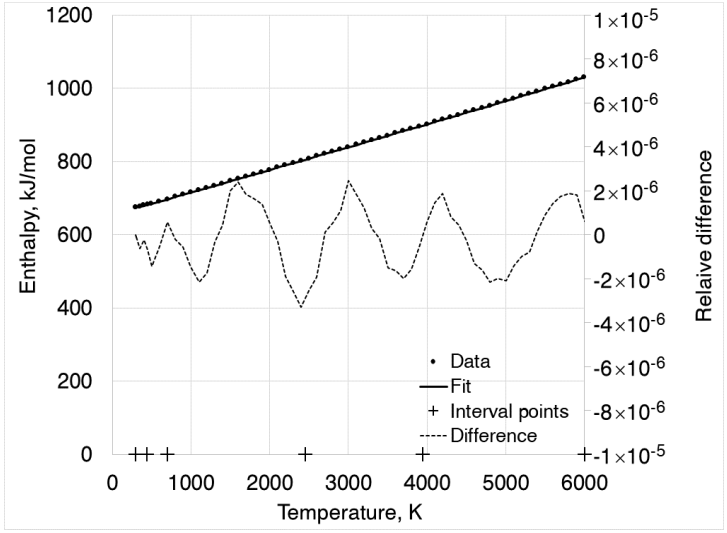


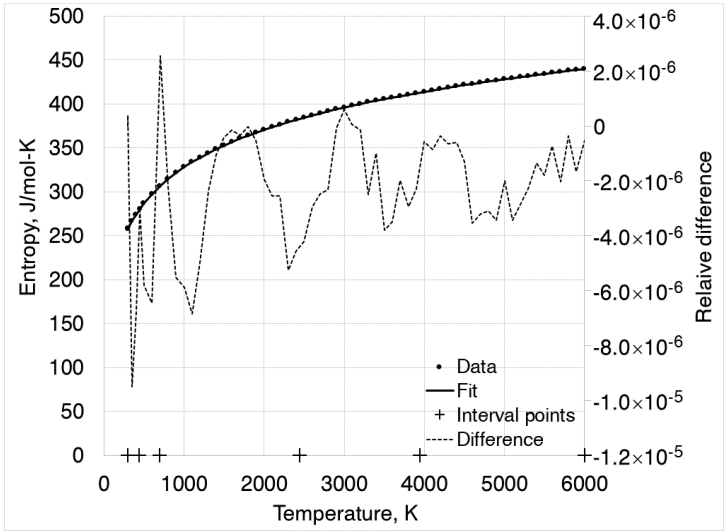


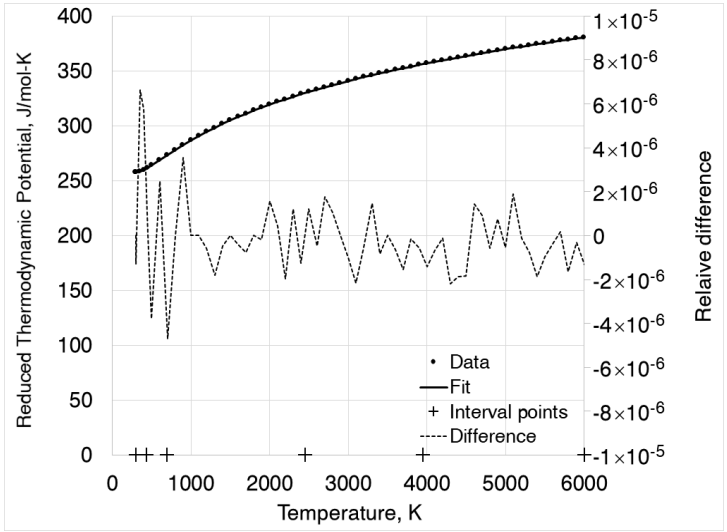
Dichloroborane, Ion (BCl_2^+) – B1Cl2+(g)

MW = 81.72 g/mol – $\Delta_f H^o$(298.15 K) = 674.042 kJ/mol – S^o(298.15 K) = 257.805 J/mol-K – p^o = 0.1 MPa



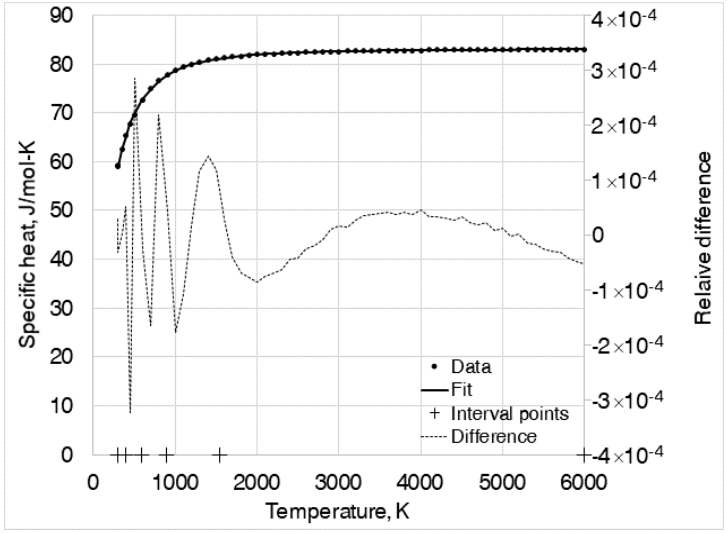


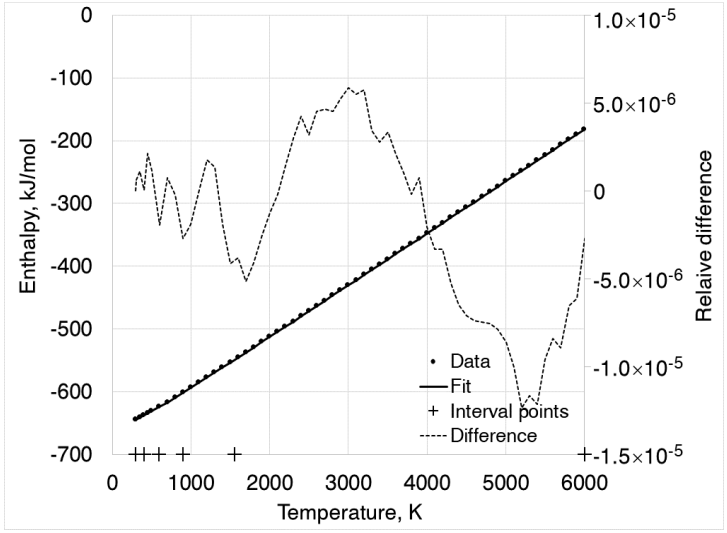


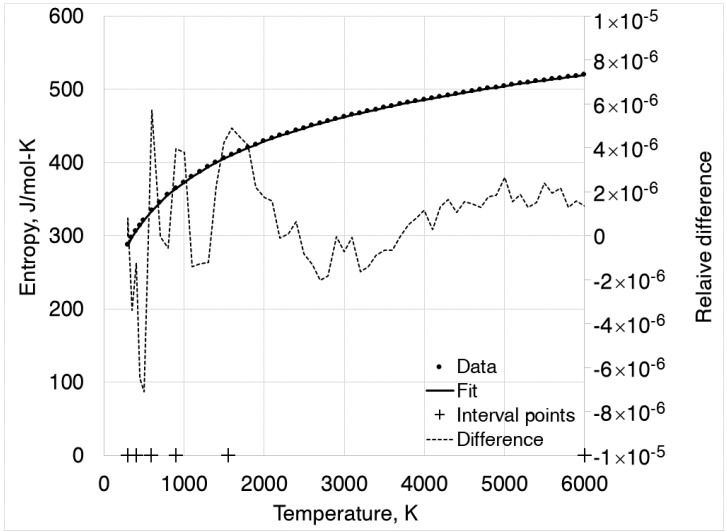


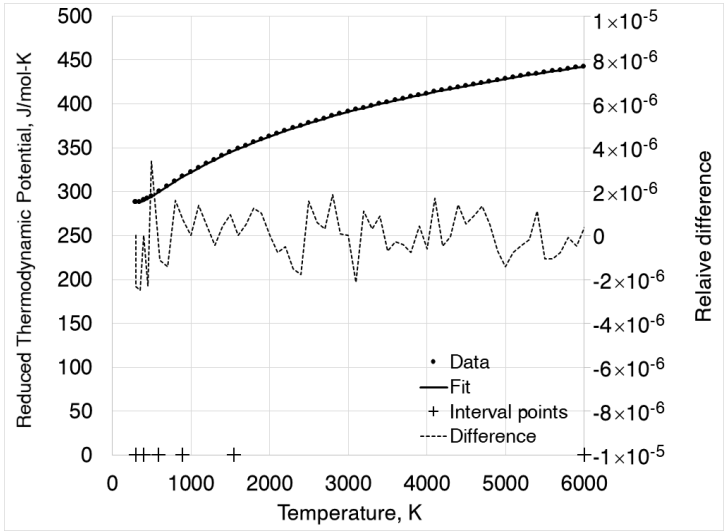
Dichlorofluoroborane (BCl_2F) – B1Cl2F1(g)

MW = 100.71 g/mol – $\Delta_f H^o$(298.15 K) = -644.336 kJ/mol – S^o(298.15 K) = 287.675 J/mol-K – p^o = 0.1 MPa



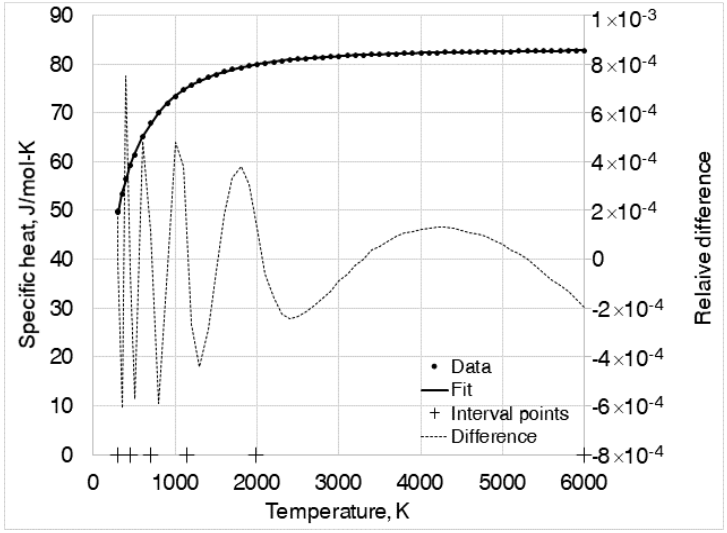


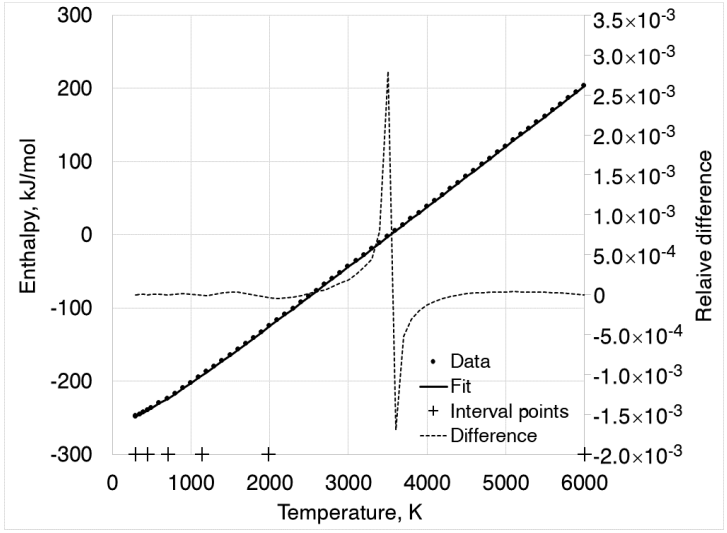


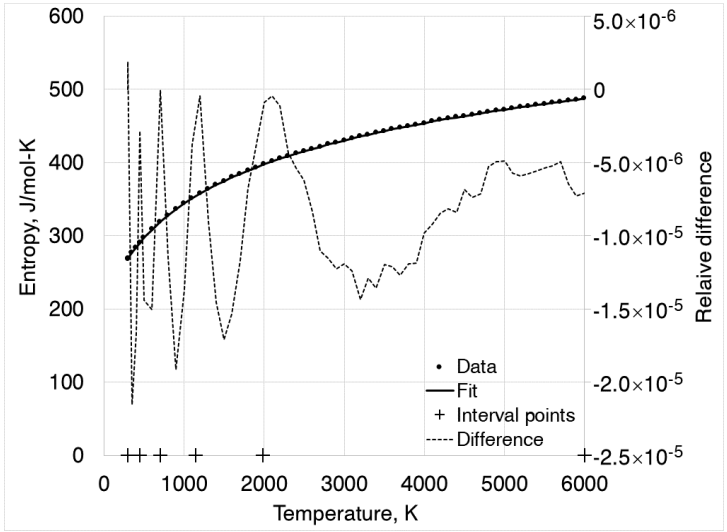


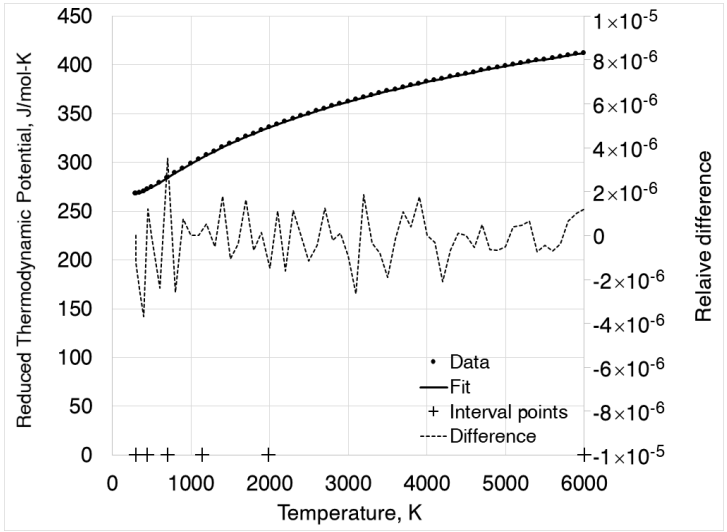
Dichloroborane ($BHCl_2$) – B1Cl2H1(g)

MW = 82.72 g/mol – $\Delta_f H^o$(298.15 K) = -248.111 kJ/mol – S^o(298.15 K) = 268.272 J/mol-K – p^o = 0.1 MPa



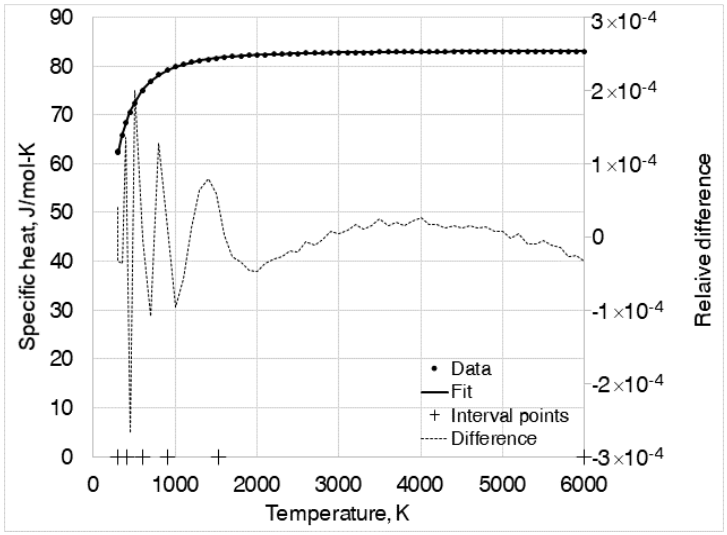


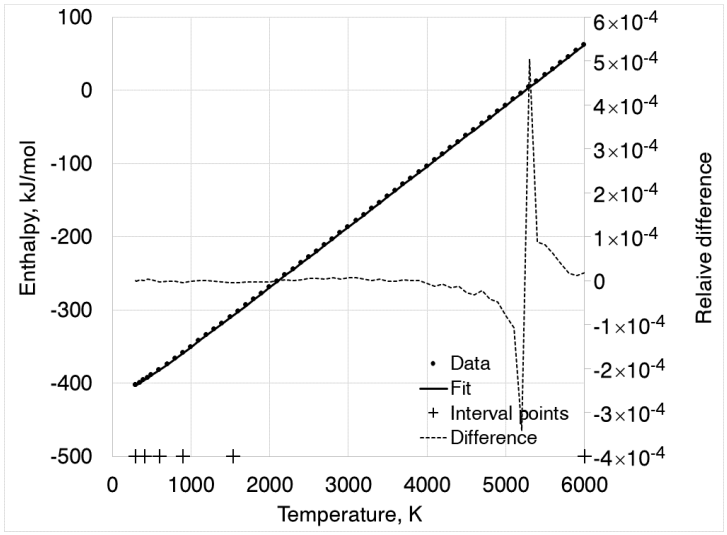


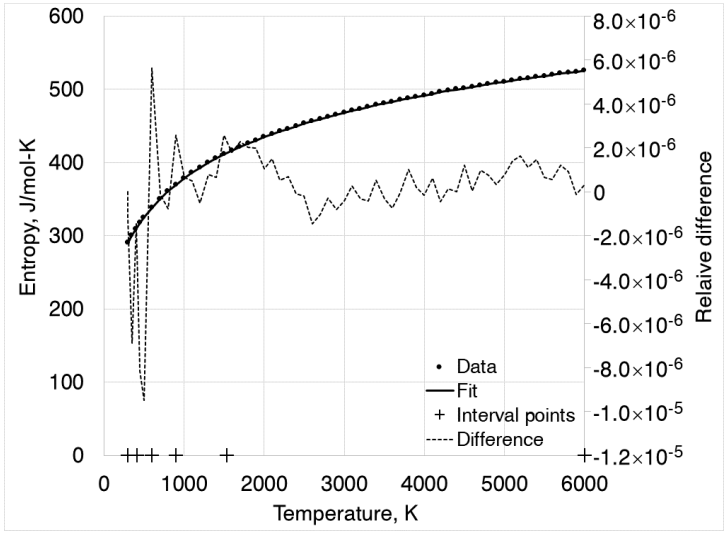


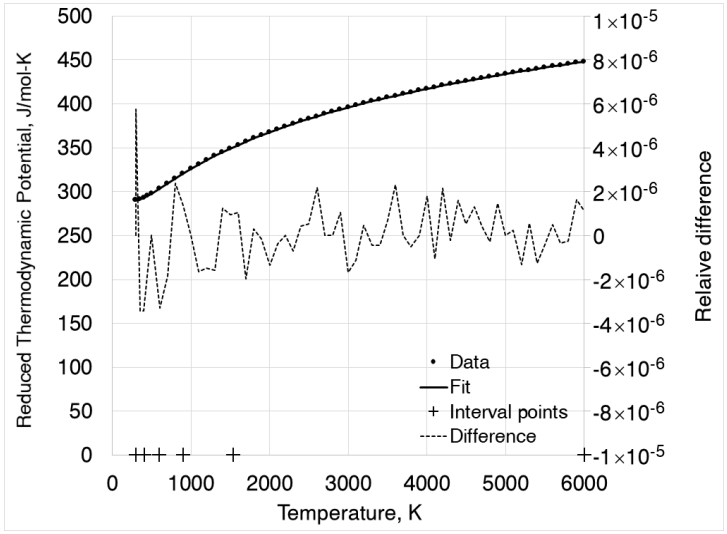
Trichloroborane (BCl_3) – B1Cl3(g)

MW = 117.17 g/mol – $\Delta_f H^o$(298.15 K) = -402.961 kJ/mol – S^o(298.15 K) = 290.169 J/mol-K – p^o = 0.1 MPa



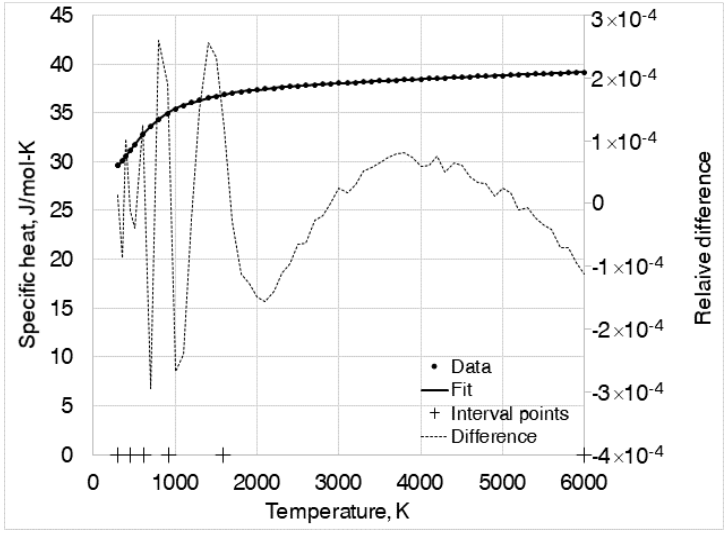


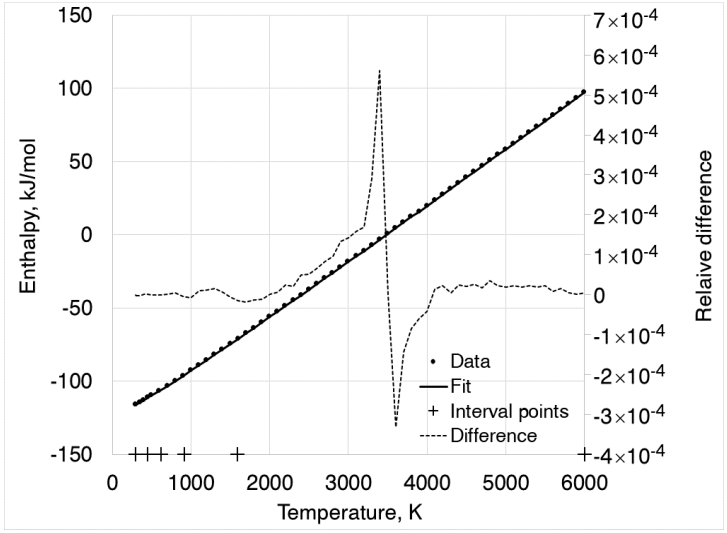


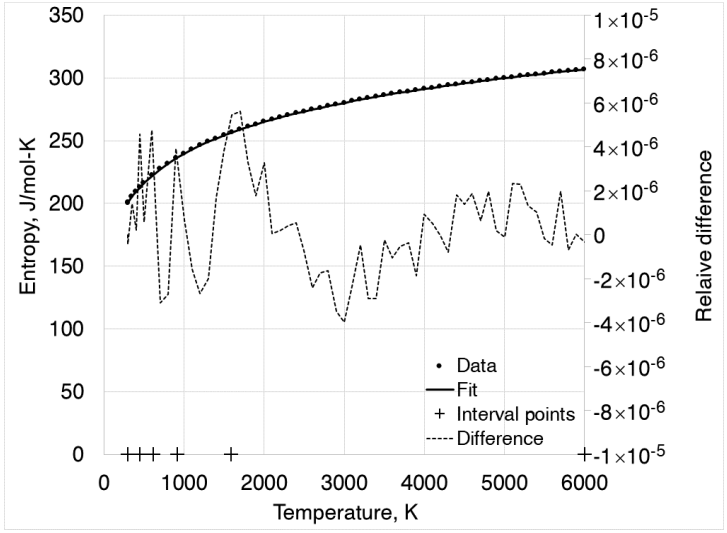


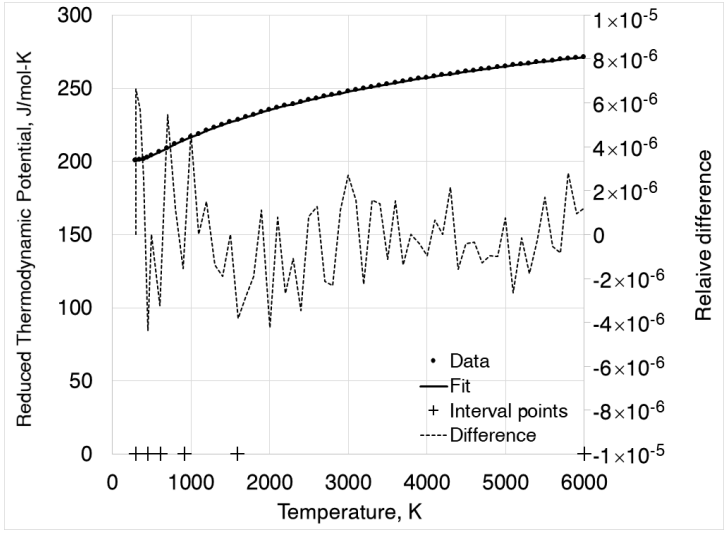
Fluoroborane (BF) – B1F1(g)

MW = 29.81 g/mol – $\Delta_f H^o$(298.15 K) = -115.897 kJ/mol – S^o(298.15 K) = 200.479 J/mol-K – p^o = 0.1 MPa



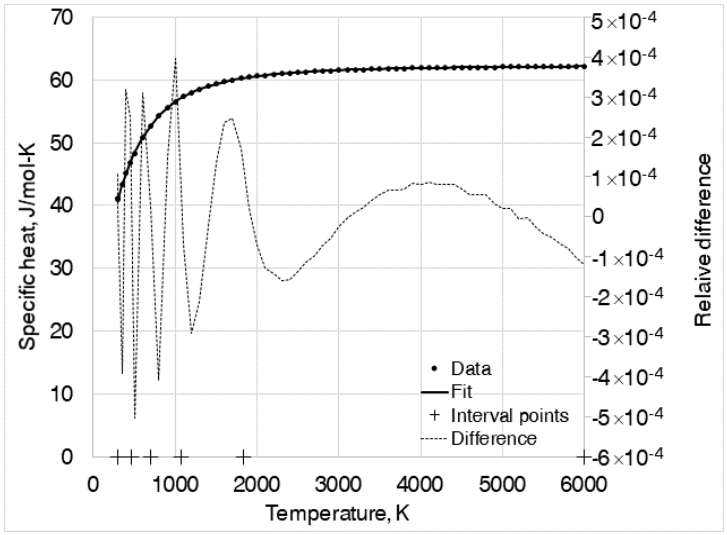


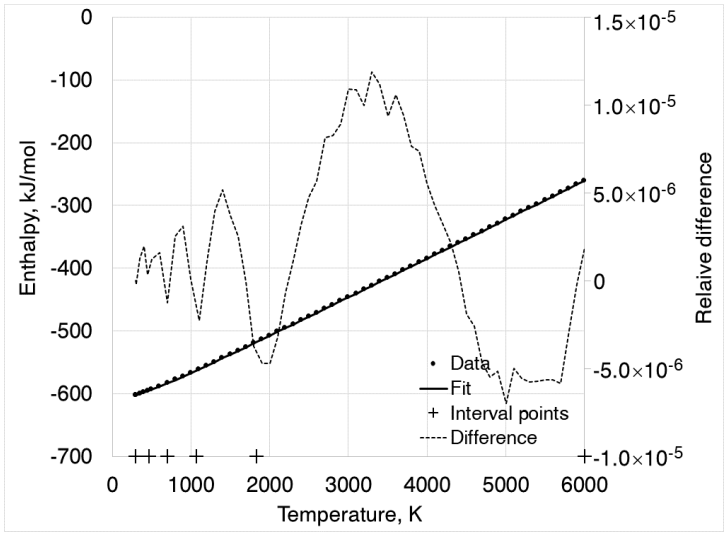


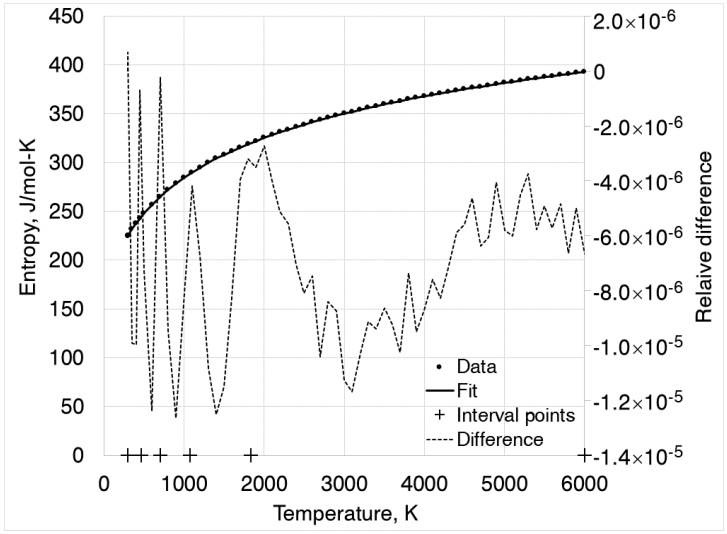


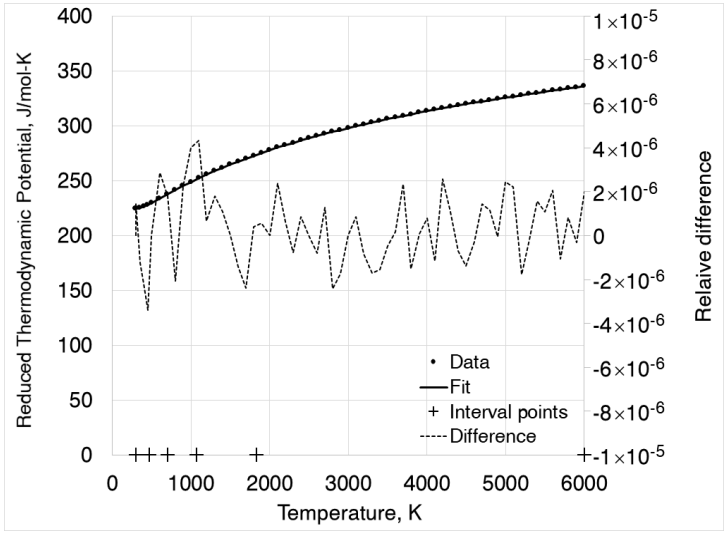
Boron Fluoride Oxide (OBF) – B1F1O1(g)

MW = 45.81 g/mol – $\Delta_f H^o$(298.15 K) = -602.496 kJ/mol – S^o(298.15 K) = 224.803 J/mol-K – p^o = 0.1 MPa



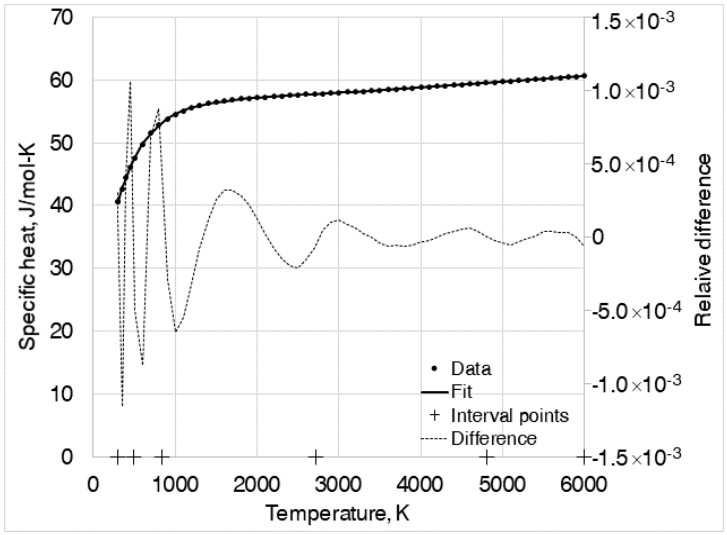


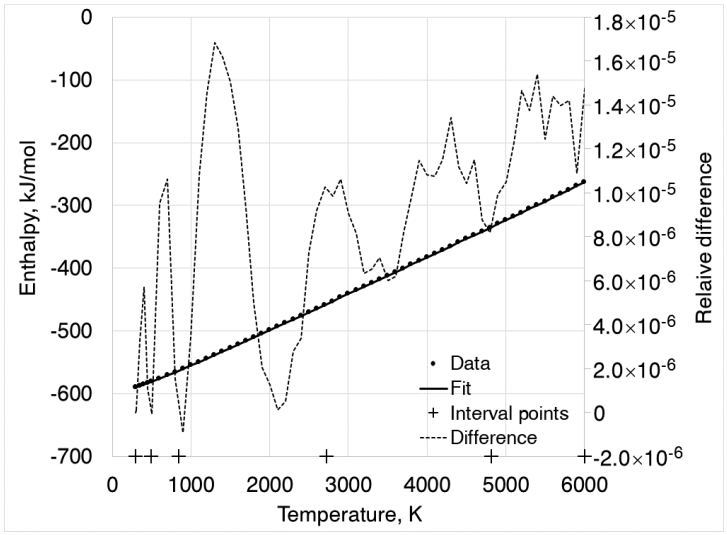


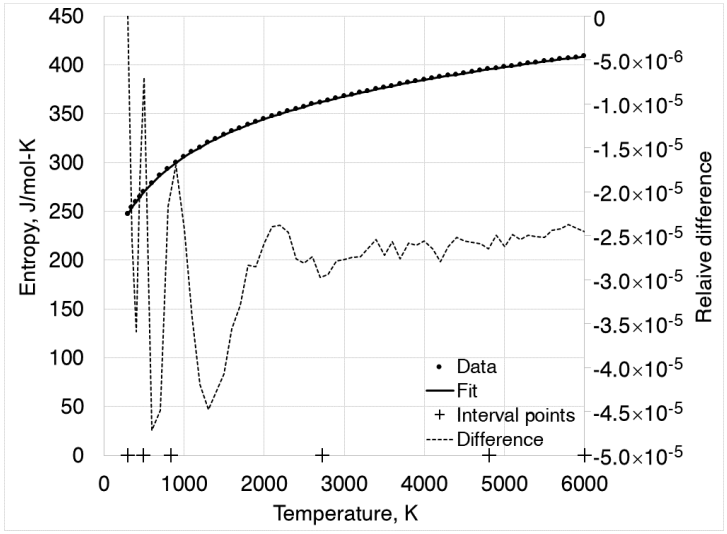


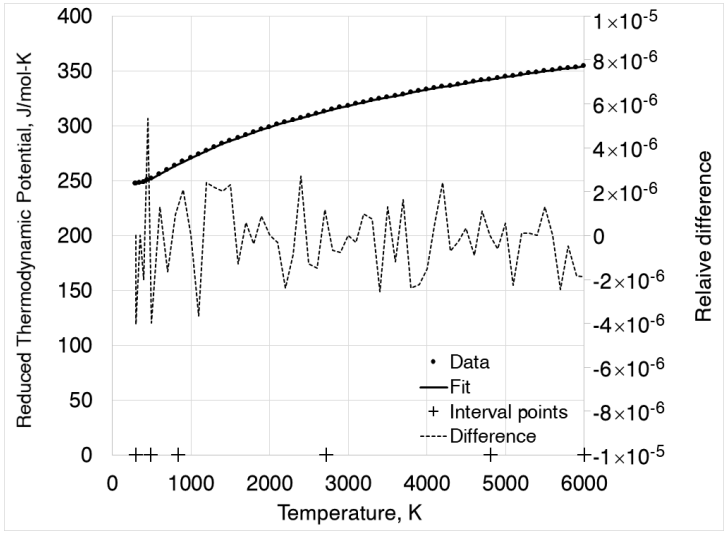
Difluoroborane (BF_2) – B1F2(g)

MW = 48.81 g/mol – $\Delta_f H^o$(298.15 K) = -589.944 kJ/mol – S^o(298.15 K) = 247.157 J/mol-K – p^o = 0.1 MPa



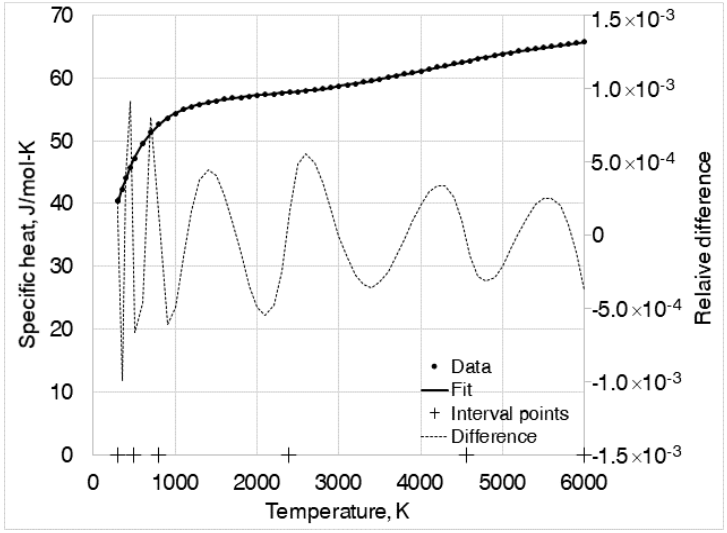


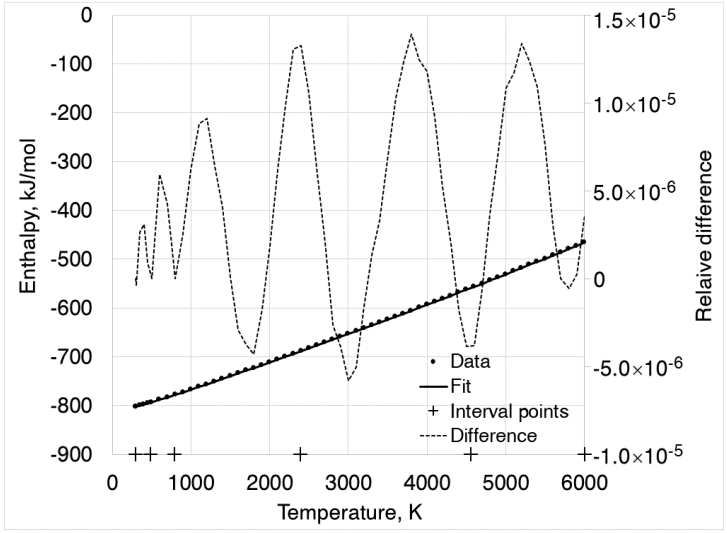


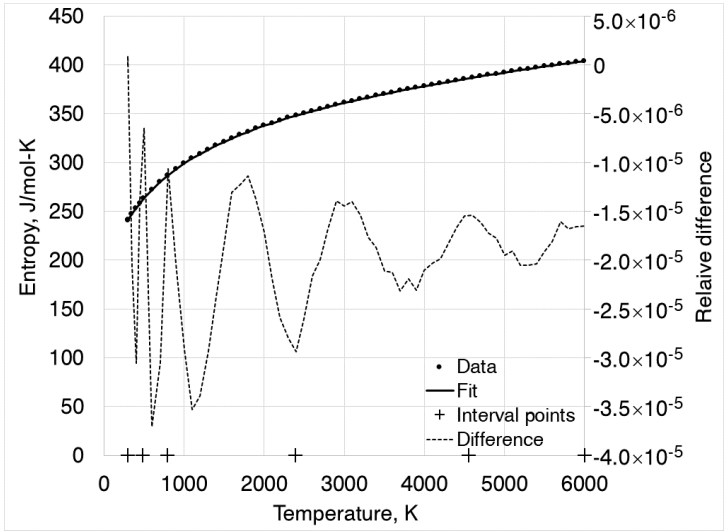


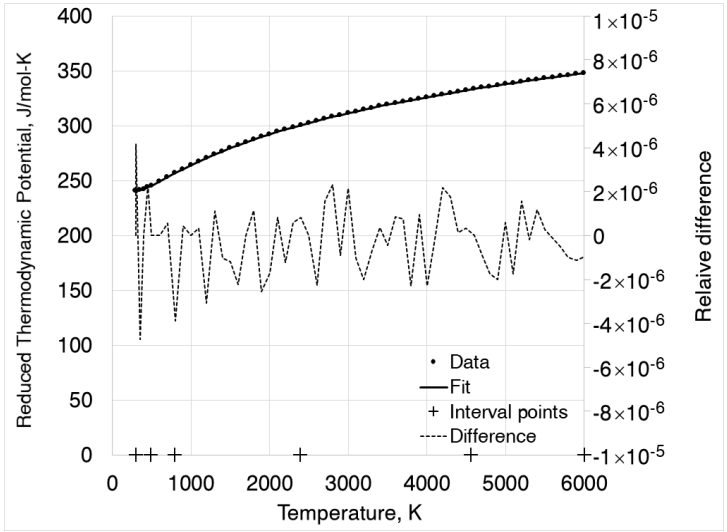
Difluoroborane, Ion (BF_2^-) – B1F2-(g)

MW = 48.81 g/mol – $\Delta_f H^o$(298.15 K) = -802.073 kJ/mol – S^o(298.15 K) = 240.74 J/mol-K – p^o = 0.1 MPa



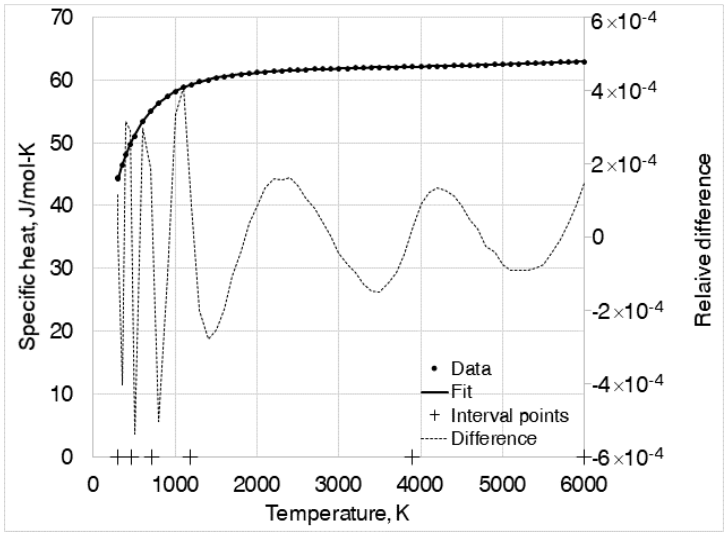


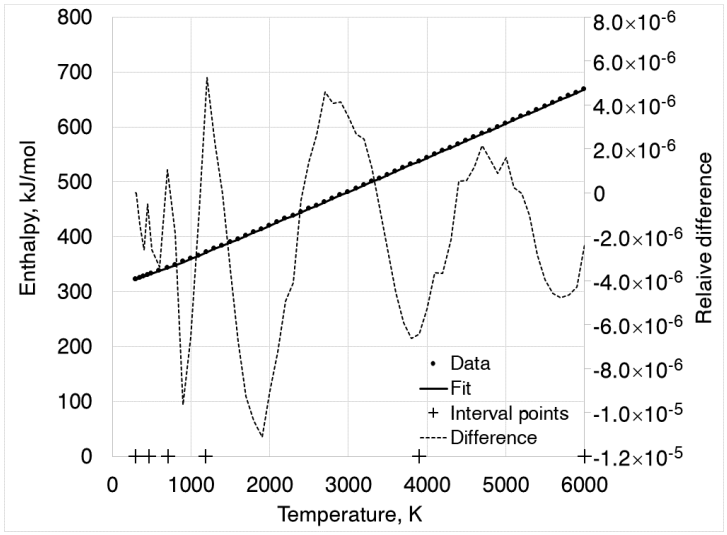


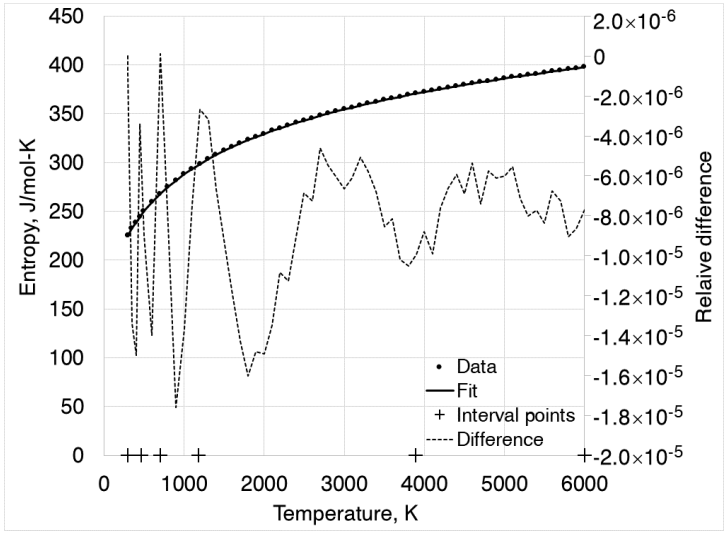


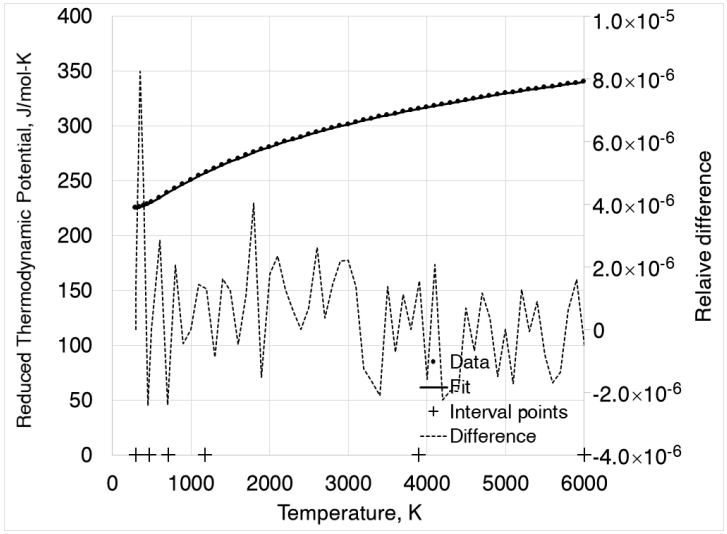
Difluoroborane, Ion (BF_2^+) – B1F2+(g)

MW = 48.81 g/mol – $\Delta_f H^o$(298.15 K) = 322.586 kJ/mol – S^o(298.15 K) = 225.148 J/mol-K – p^o = 0.1 MPa



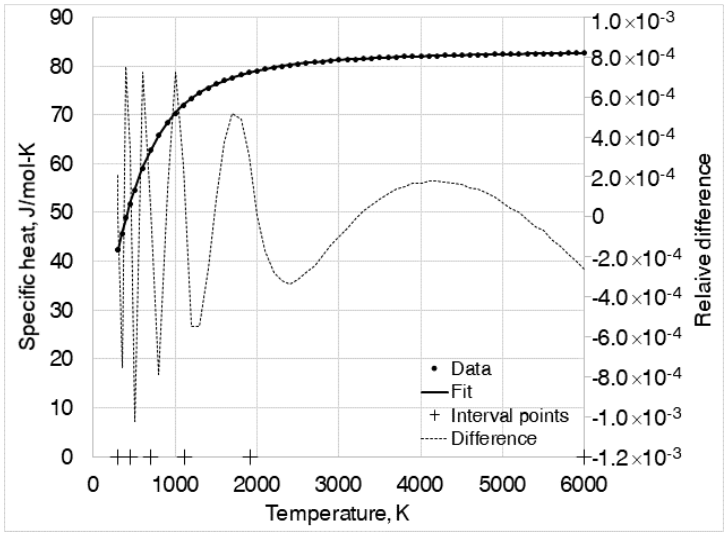


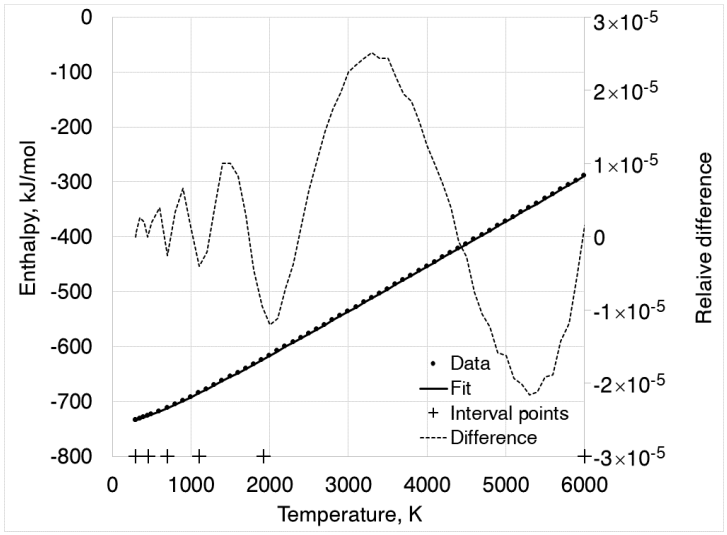


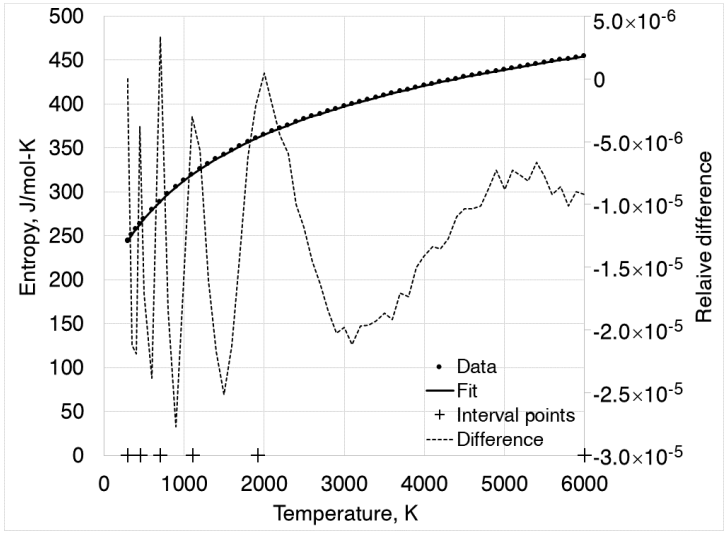


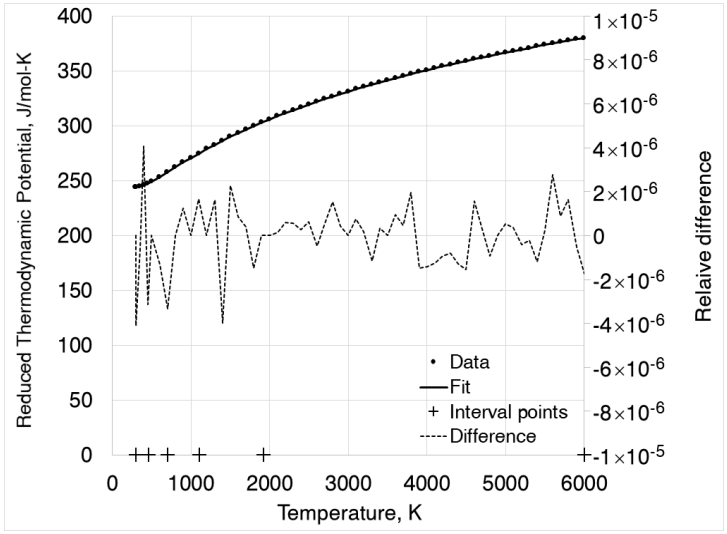
Difluoroborane (BHF_2) – B1F2H1(g)

MW = 49.81 g/mol – $\Delta_f H^o$(298.15 K) = -733.874 kJ/mol – S^o(298.15 K) = 244.029 J/mol-K – p^o = 0.1 MPa



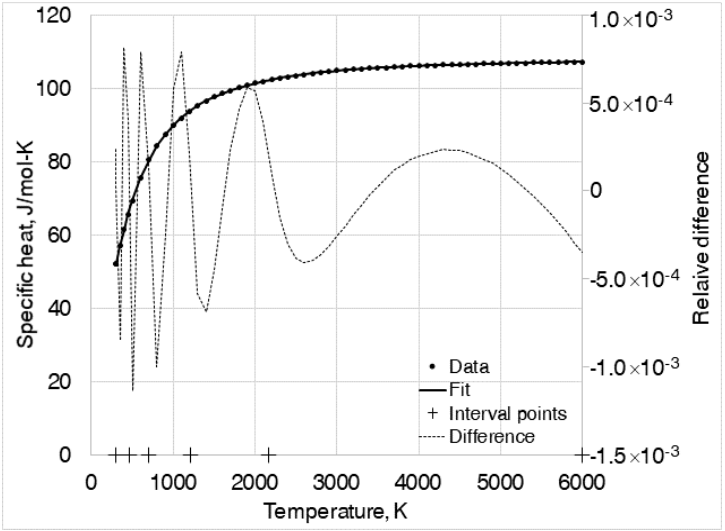


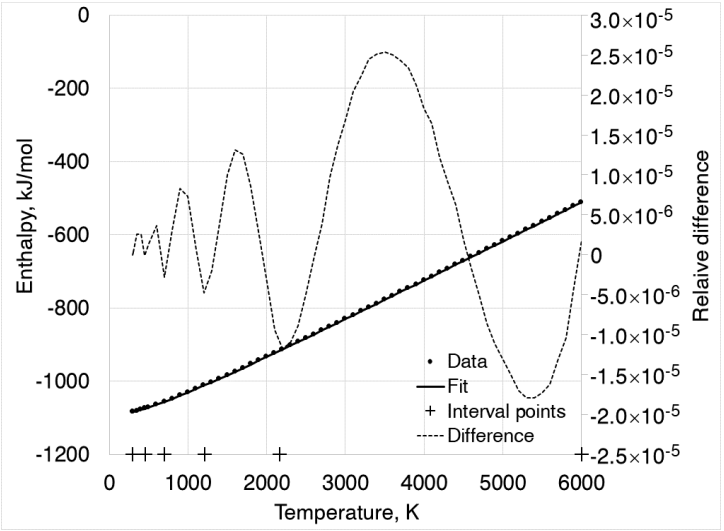


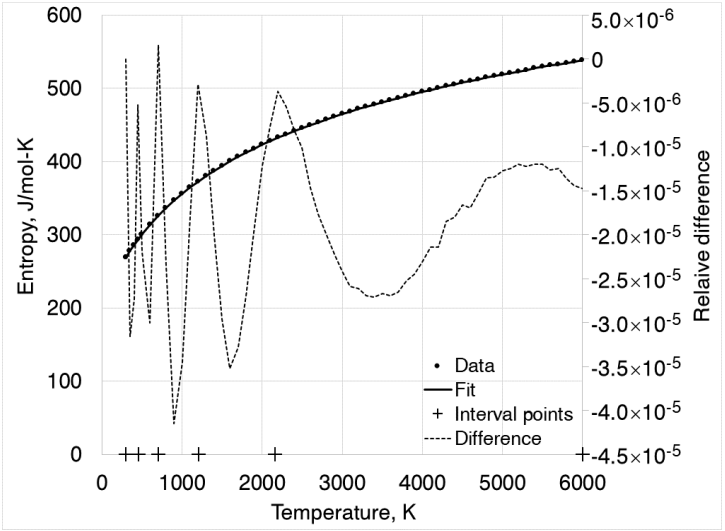


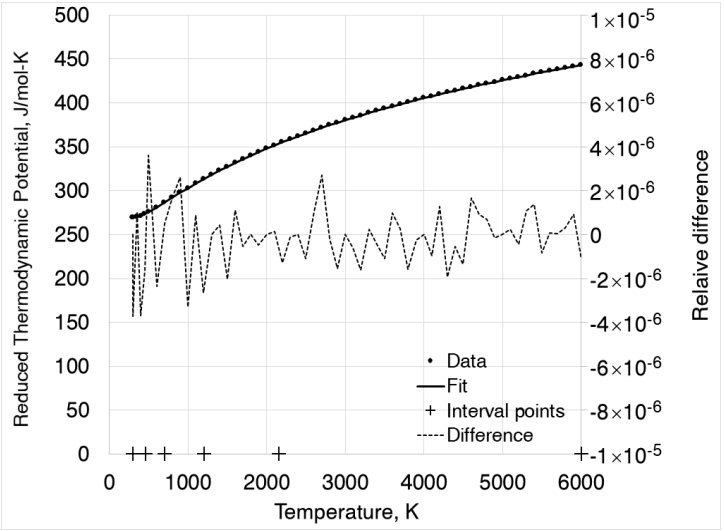
Difluorohydroxyborane (BF_2OH) – B1F2H1O1(g)

MW = 65.81 g/mol – $\Delta_f H^o$(298.15 K) = -1083.656 kJ/mol – S^o(298.15 K) = 269.174 J/mol-K – p^o = 0.1 MPa



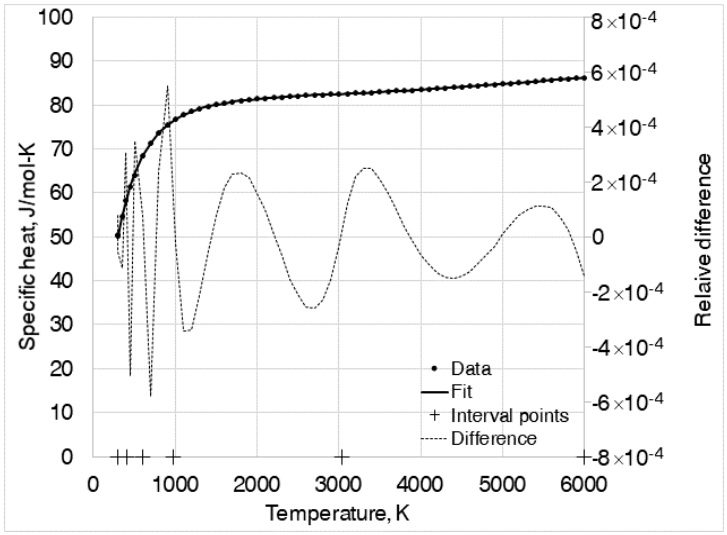


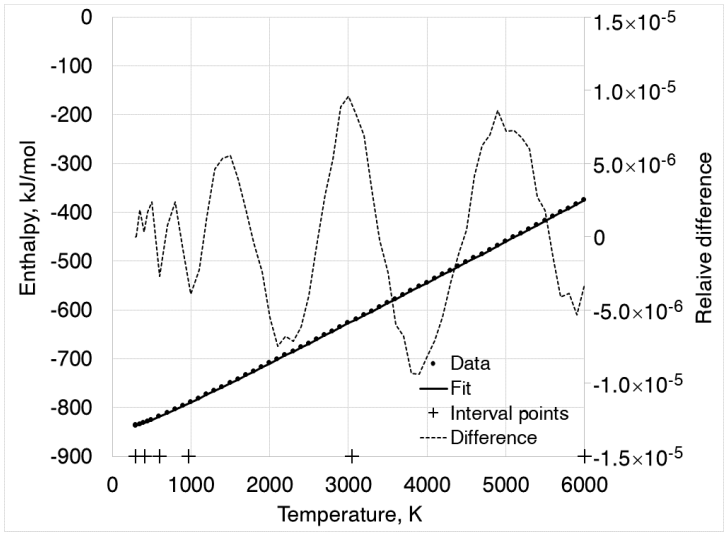


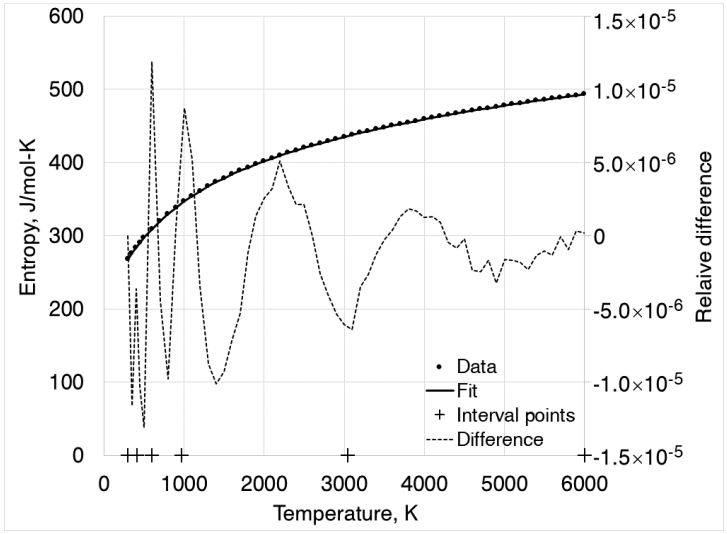


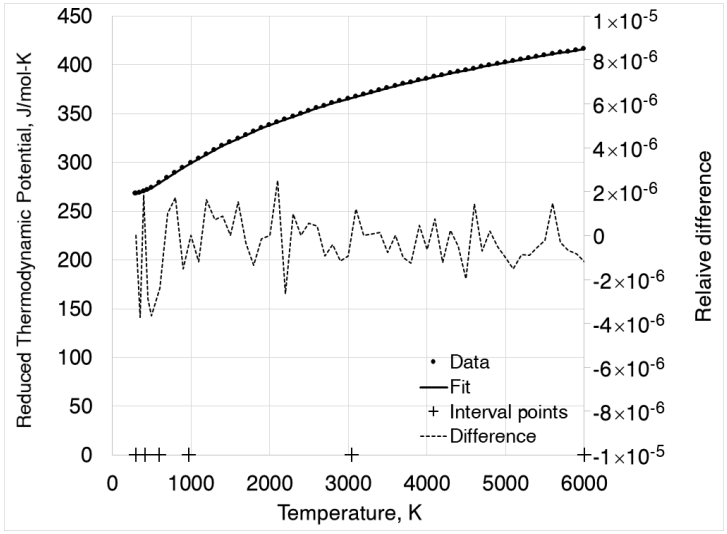
Boron Fluoride Oxide (OBF_2) – B1F2O1(g)

MW = 64.81 g/mol – $\Delta_f H^o$(298.15 K) = -836.8 kJ/mol – S^o(298.15 K) = 267.842 J/mol-K – p^o = 0.1 MPa



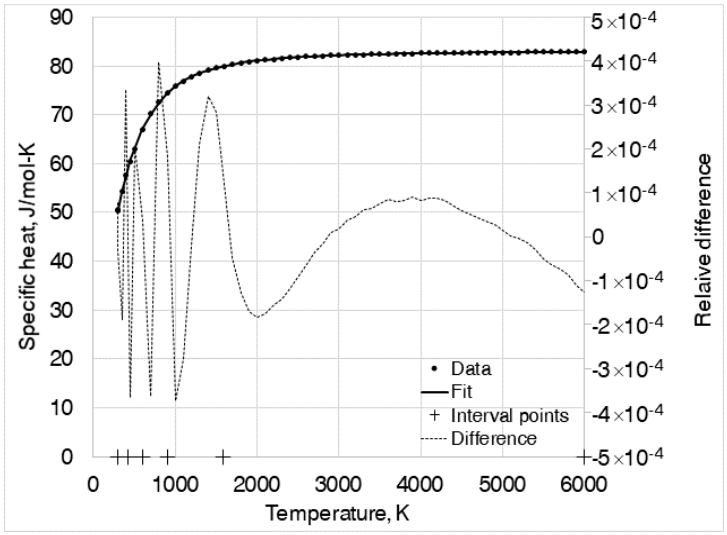


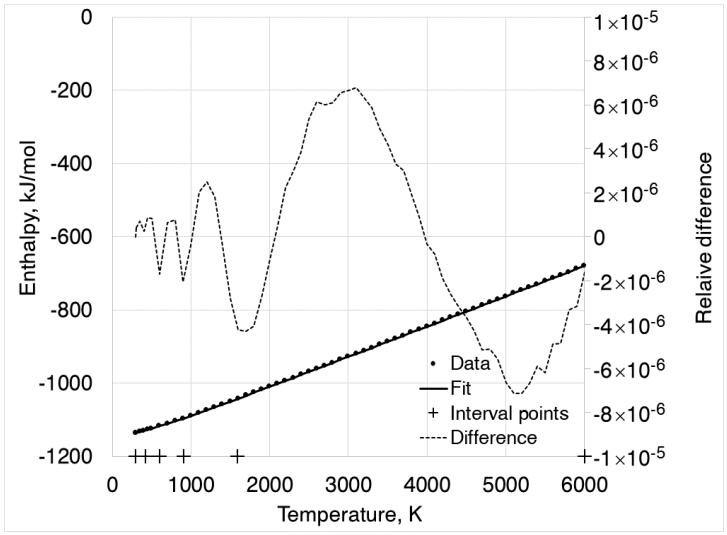


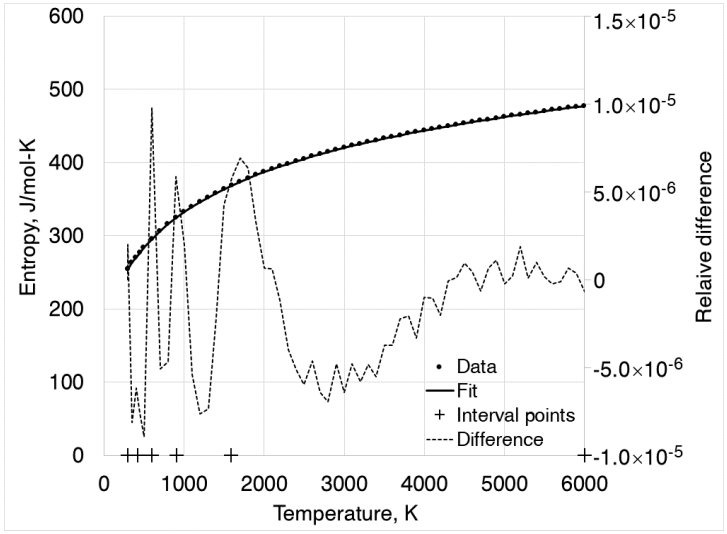


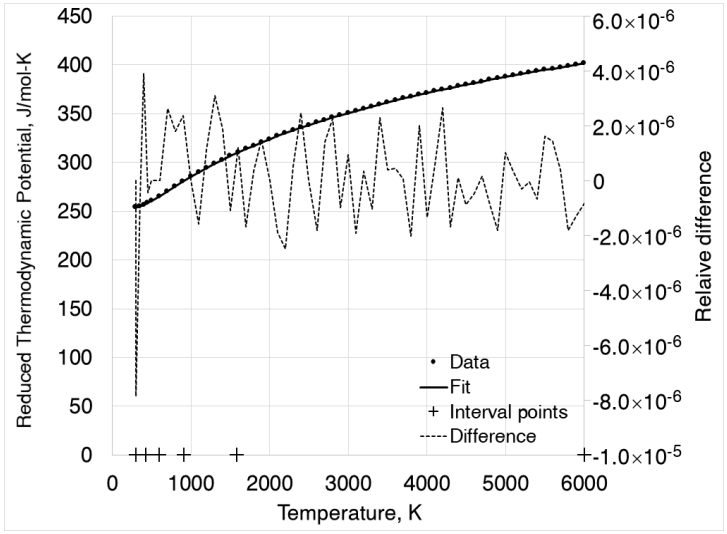
Trifluoroborane (BF_3) – B1F3(g)

MW = 67.81 g/mol – $\Delta_f H^o$(298.15 K) = -1135.621 kJ/mol – S^o(298.15 K) = 254.355 J/mol-K – p^o = 0.1 MPa



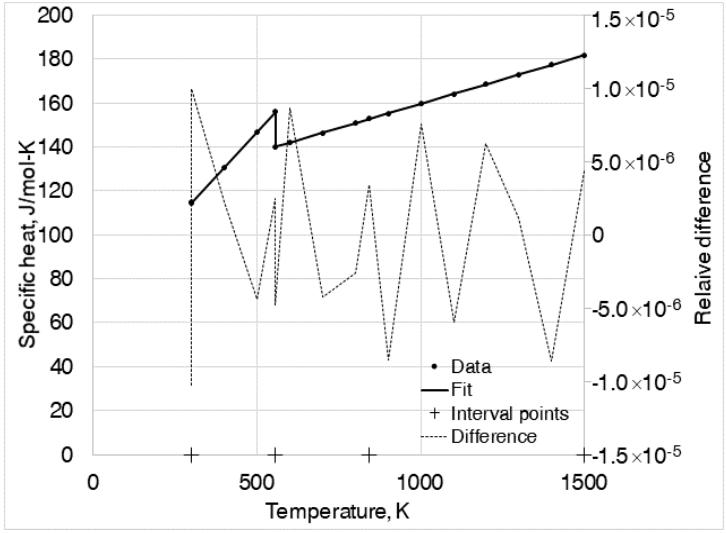


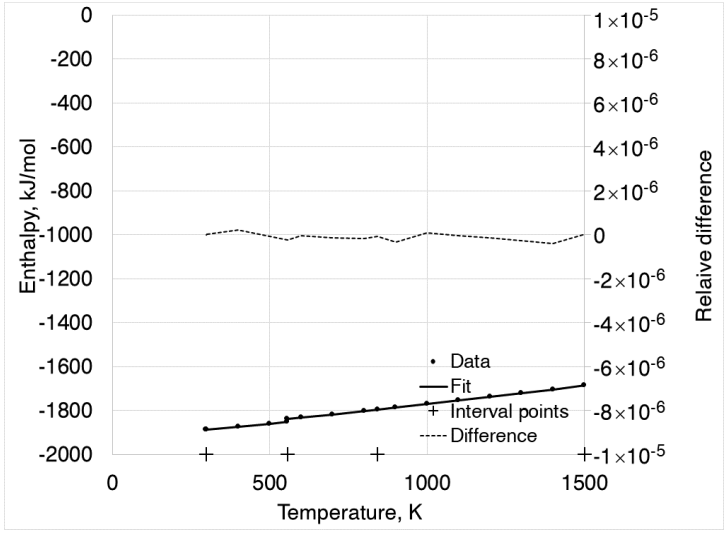


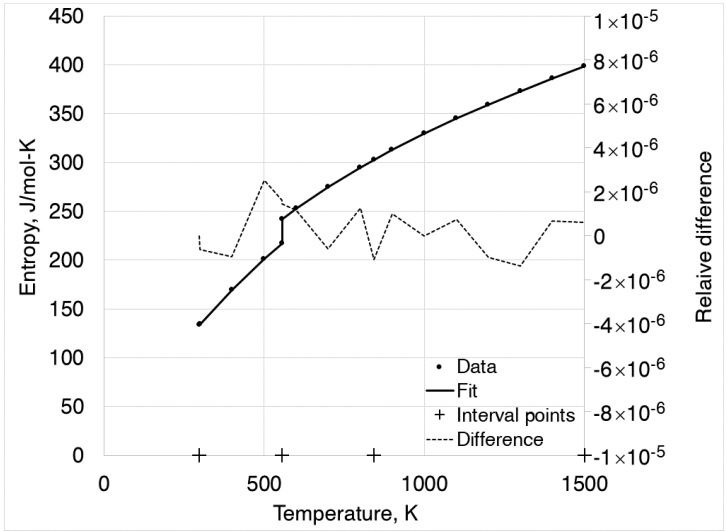


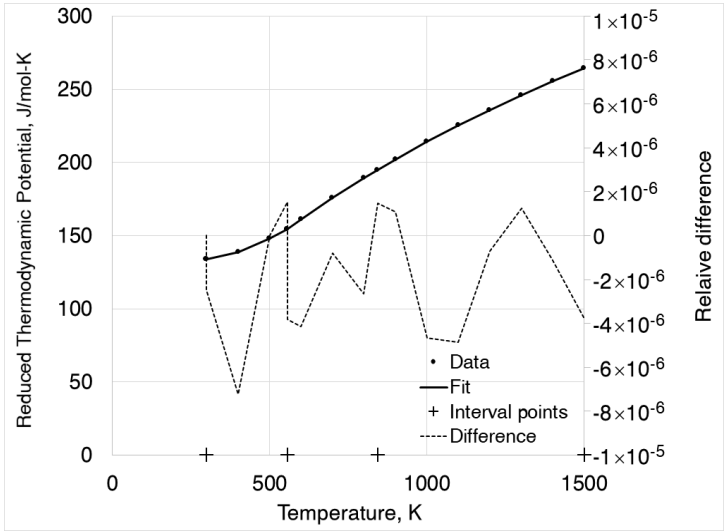
Potassium Tetrafluoroborate (KBF_4) – B1F4K1(cr)

MW = 125.9 g/mol – $\Delta_f H^o$(298.15 K) = -1886.984 kJ/mol – S^o(298.15 K) = 133.888 J/mol-K – p^o = 0.1 MPa



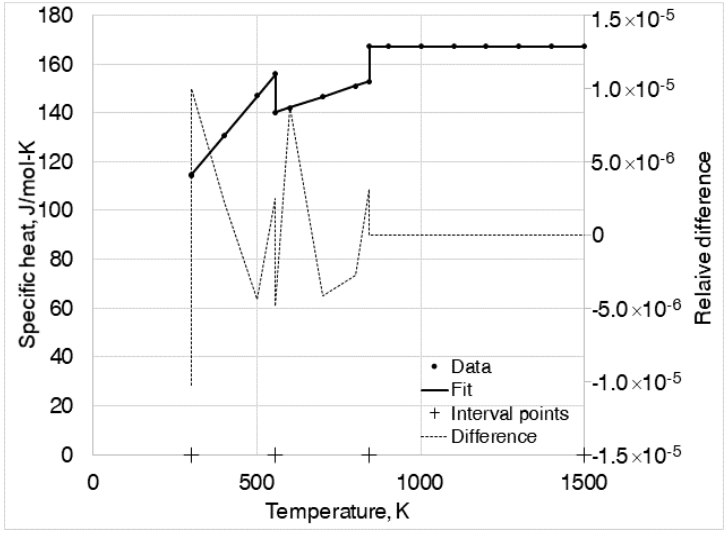


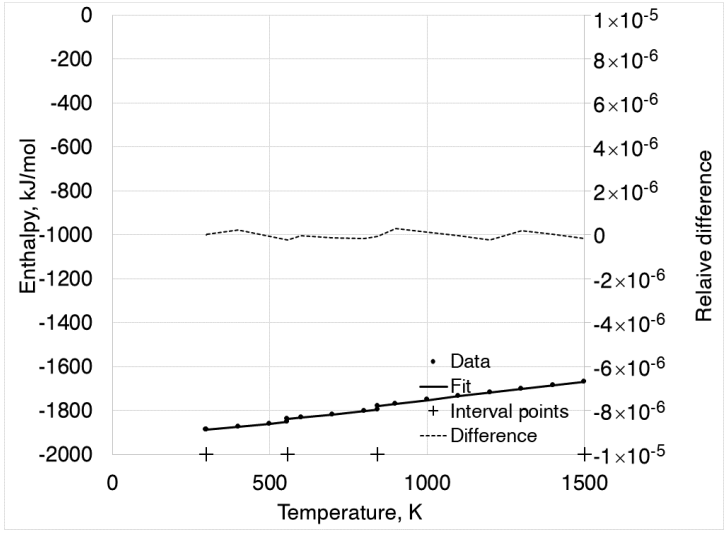


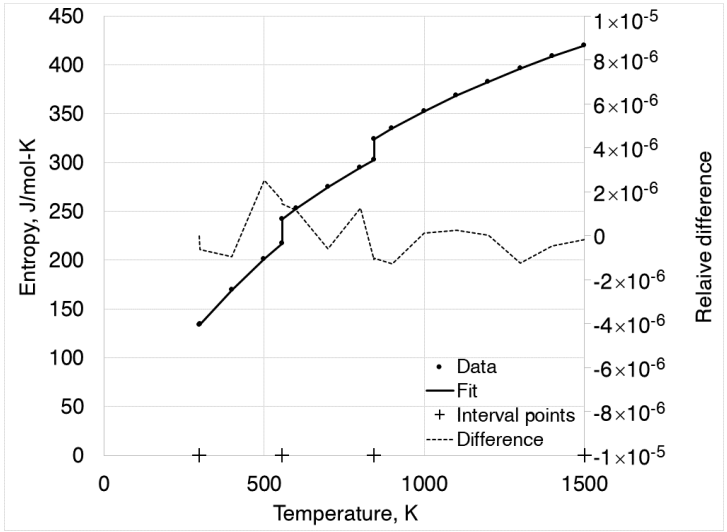


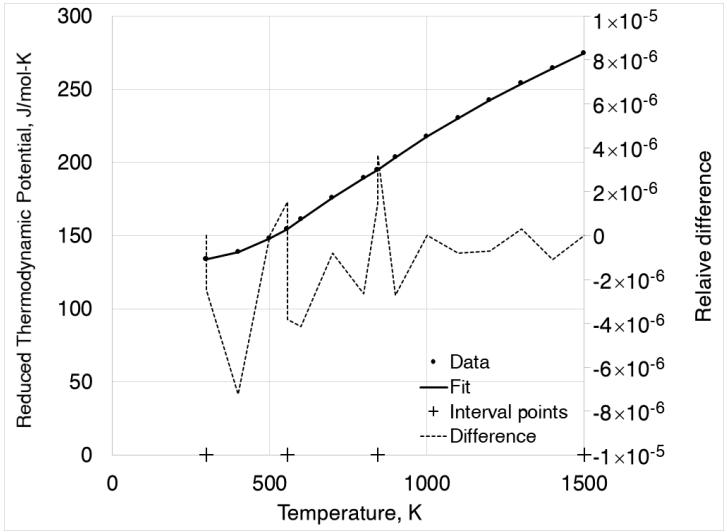
Potassium Tetrafluoroborate (KBF_4) – B1F4K1(cr,l)

MW = 125.9 g/mol – $\Delta_f H^o$(298.15 K) = -1886.984 kJ/mol – S^o(298.15 K) = 133.888 J/mol-K – p^o = 0.1 MPa



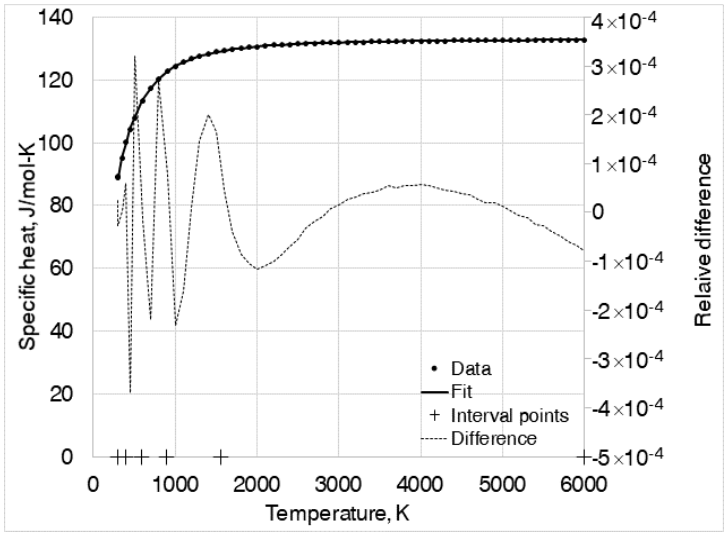


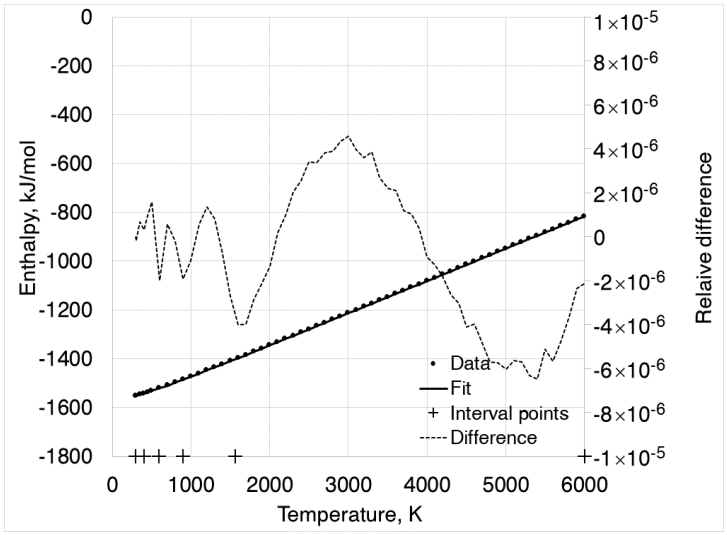


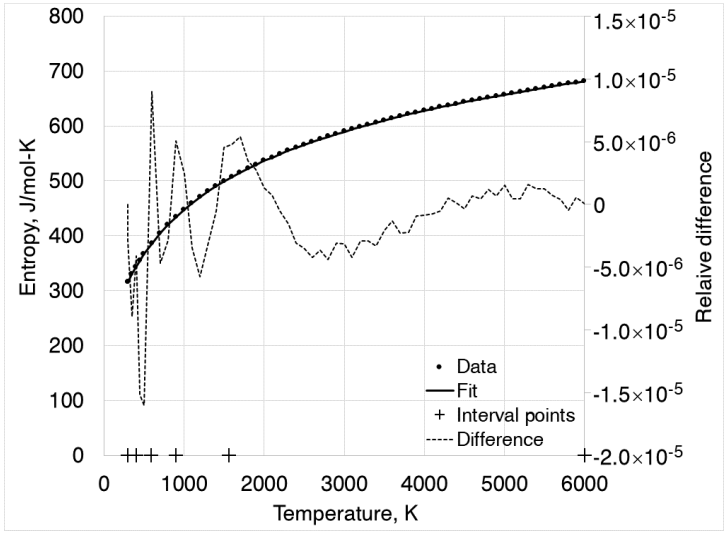


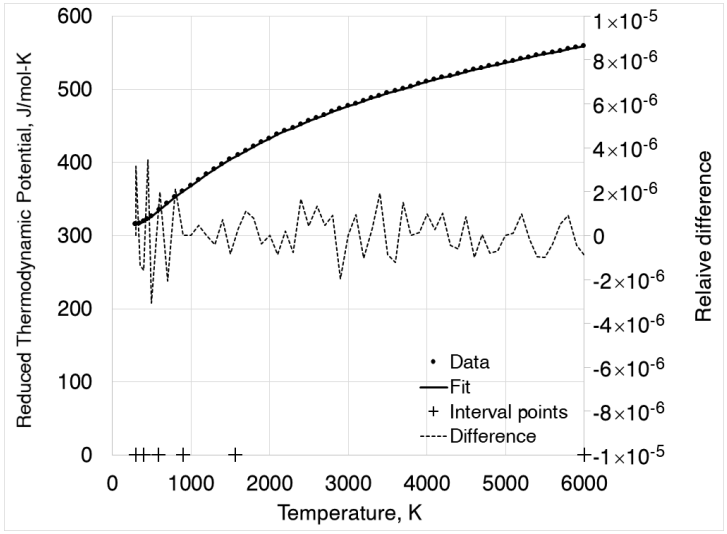
Potassium Tetrafluoroborate (KBF_4) – B1F4K1(g)

MW = 125.9 g/mol – $\Delta_f H^o$(298.15 K) = -1552.264 kJ/mol – S^o(298.15 K) = 315.394 J/mol-K – p^o = 0.1 MPa



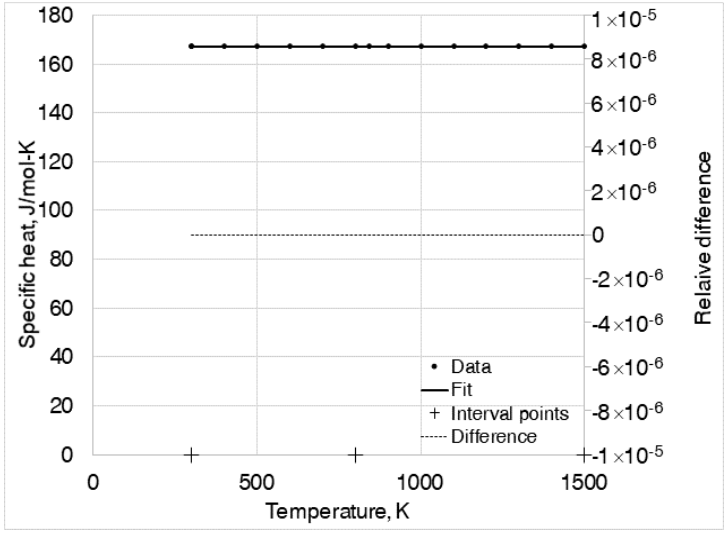


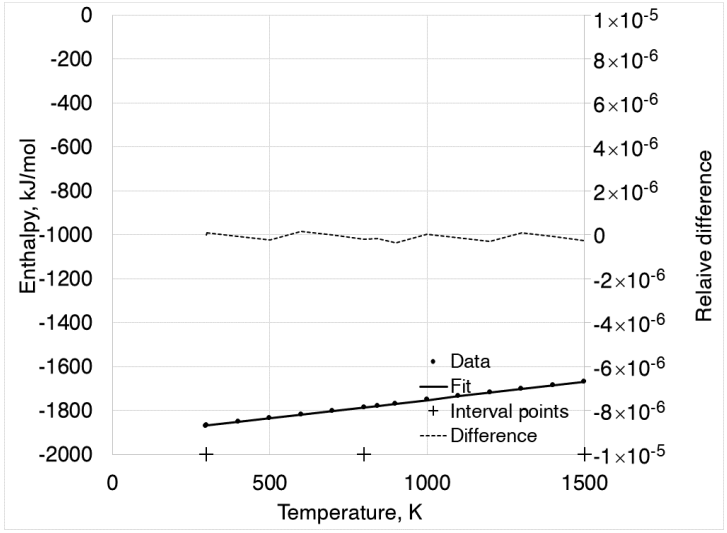


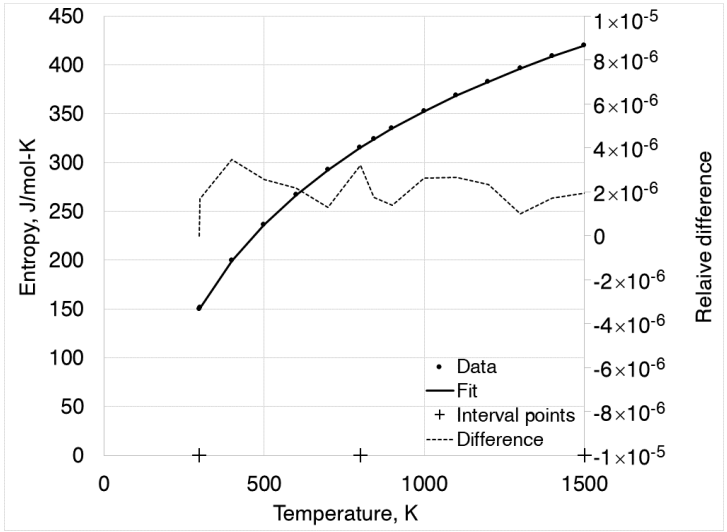


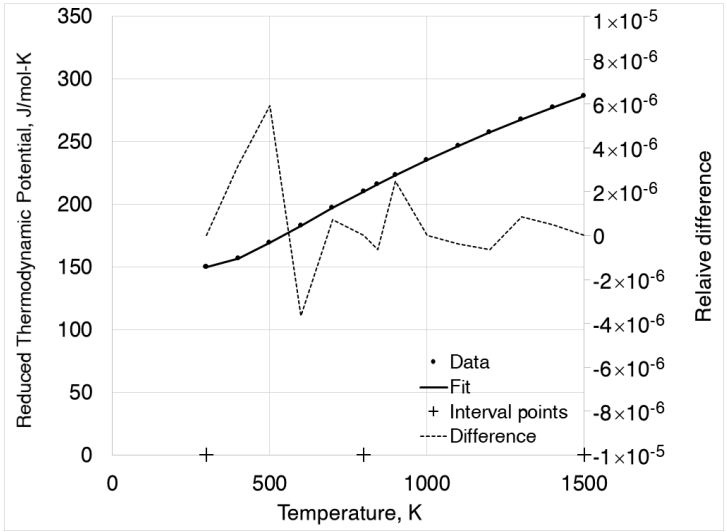
Potassium Tetrafluoroborate (KBF_4) – B1F4K1(l)

MW = 125.9 g/mol – $\Delta_f H^o$(298.15 K) = -1869.415 kJ/mol – S^o(298.15 K) = 150.129 J/mol-K – p^o = 0.1 MPa



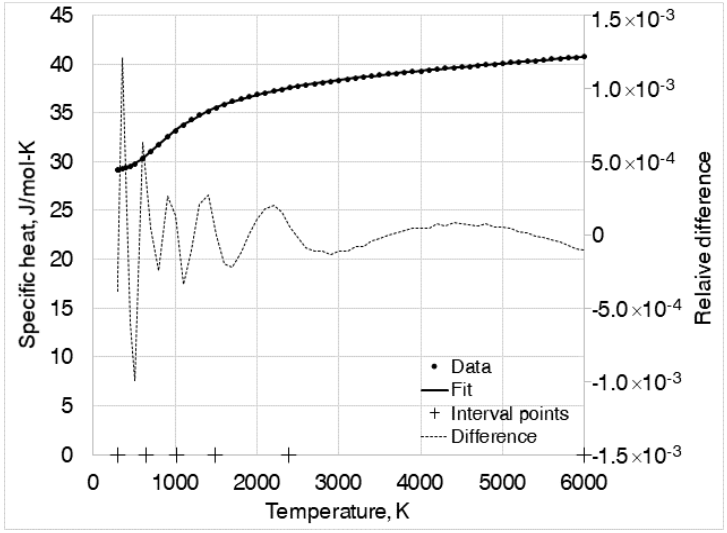


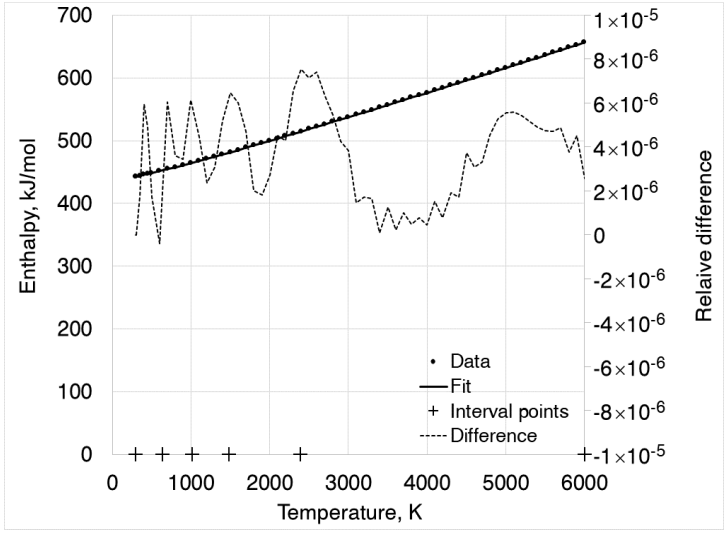


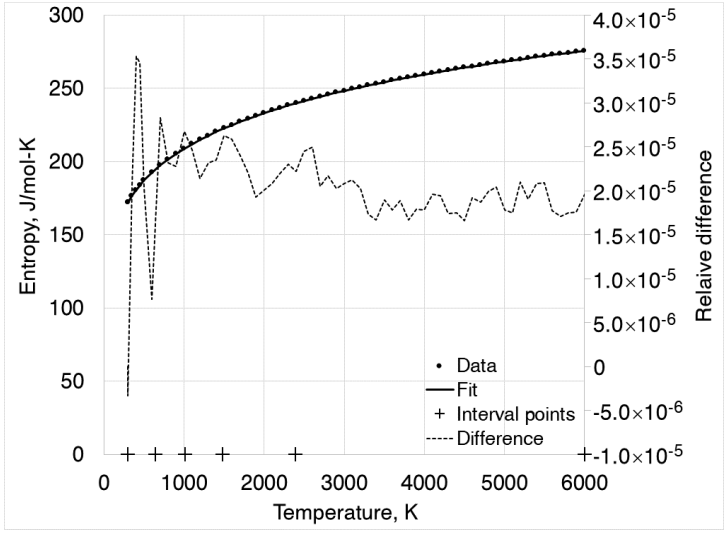


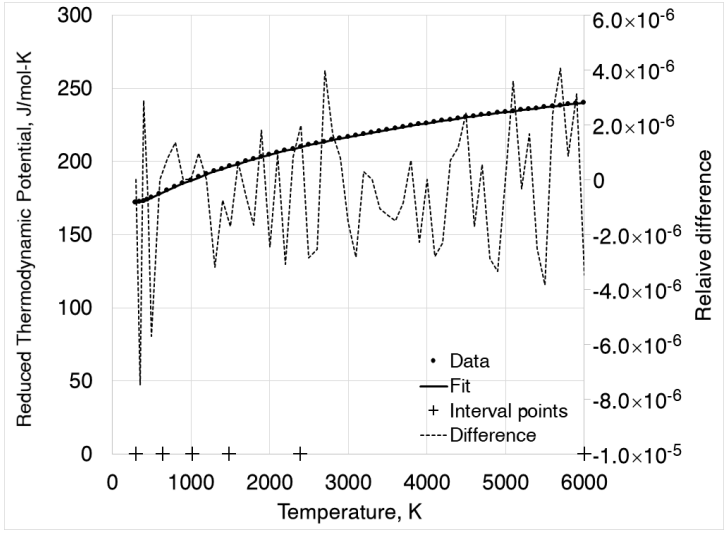
Borane (BH) – B1H1(g)

MW = 11.82 g/mol – $\Delta_f H^o$(298.15 K) = 442.667 kJ/mol – S^o(298.15 K) = 171.849 J/mol-K – p^o = 0.1 MPa



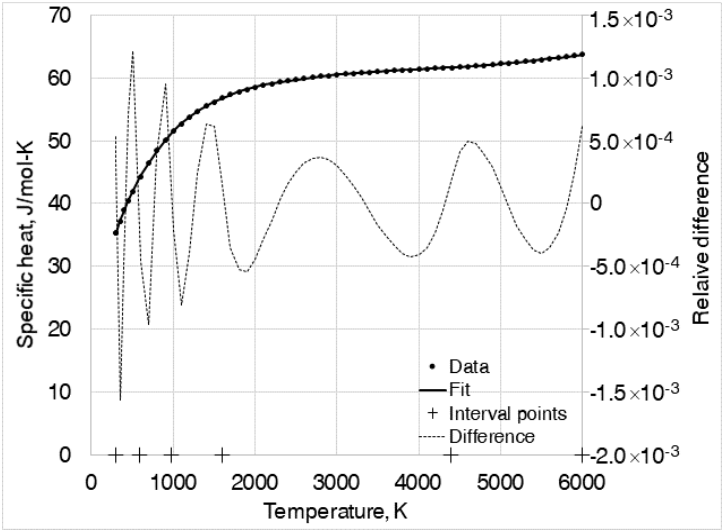


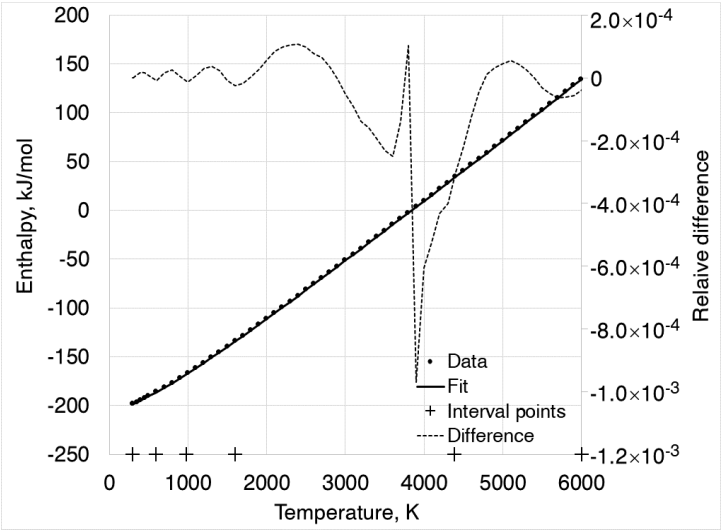


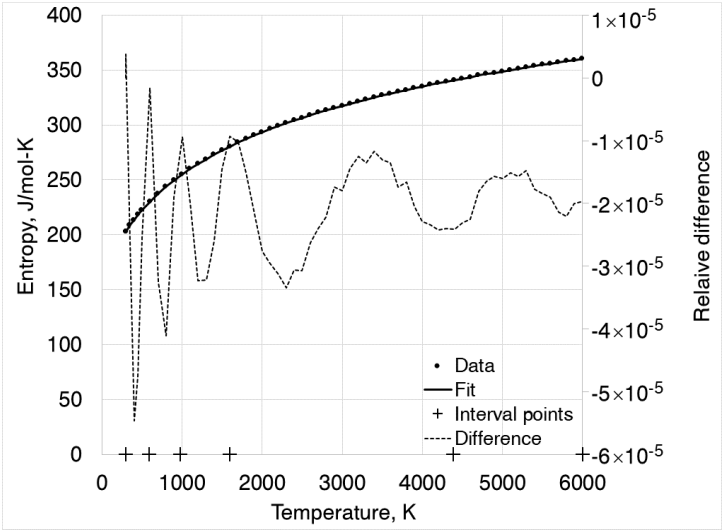


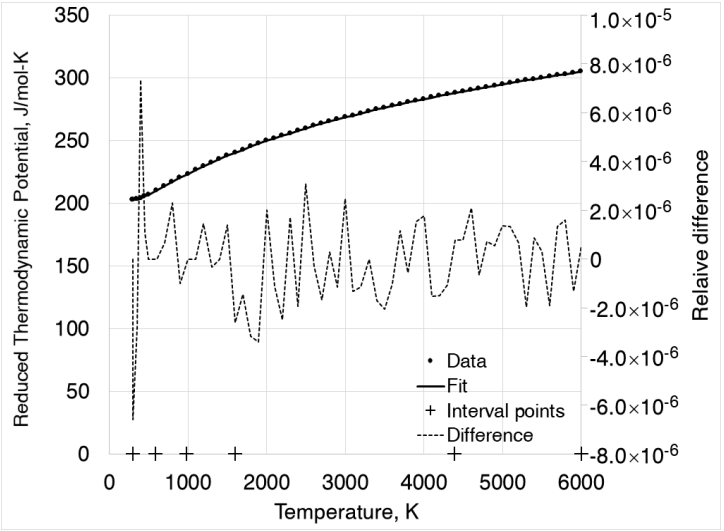
Boron Hydride Oxide (HBO) – B1H1O1(g)

MW = 27.82 g/mol – $\Delta_f H^o$(298.15 K) = -198.322 kJ/mol – S^o(298.15 K) = 202.617 J/mol-K – p^o = 0.1 MPa



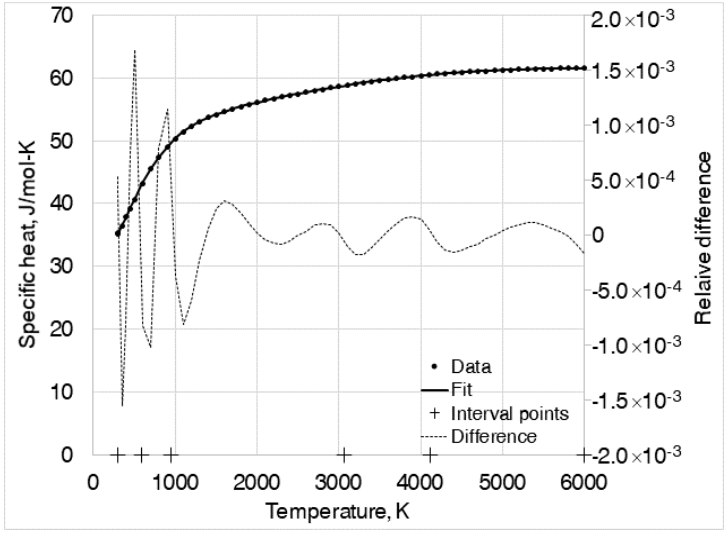


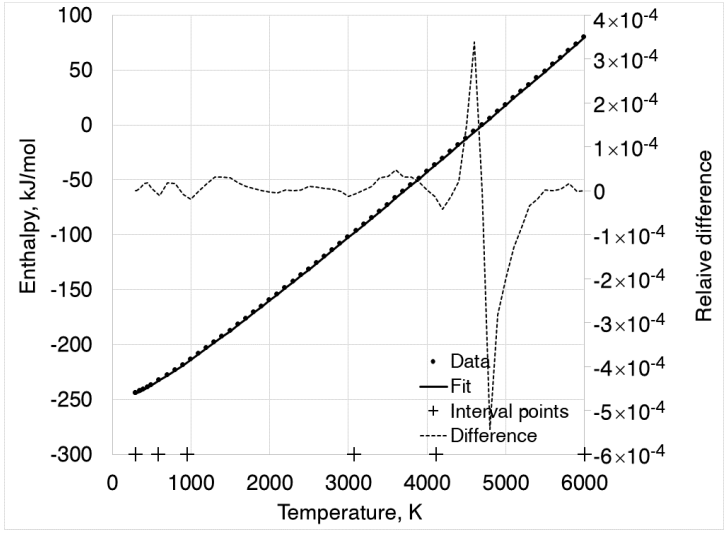


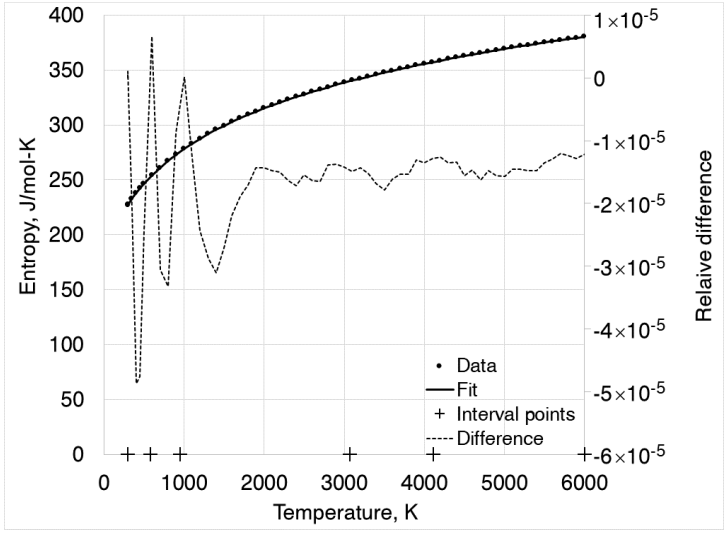


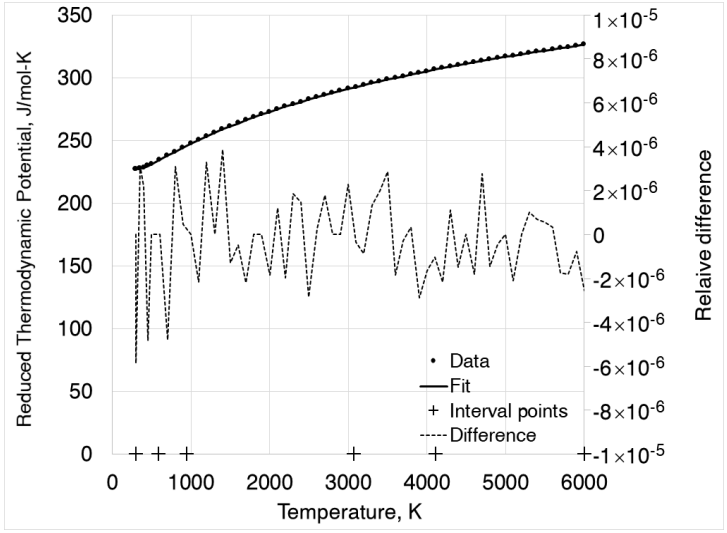
Boron Hydride Oxide, Ion (HBO^-) – B1H1O1-(g)

MW = 27.82 g/mol – $\Delta_f H^o$(298.15 K) = -244.471 kJ/mol – S^o(298.15 K) = 227.058 J/mol-K – p^o = 0.1 MPa



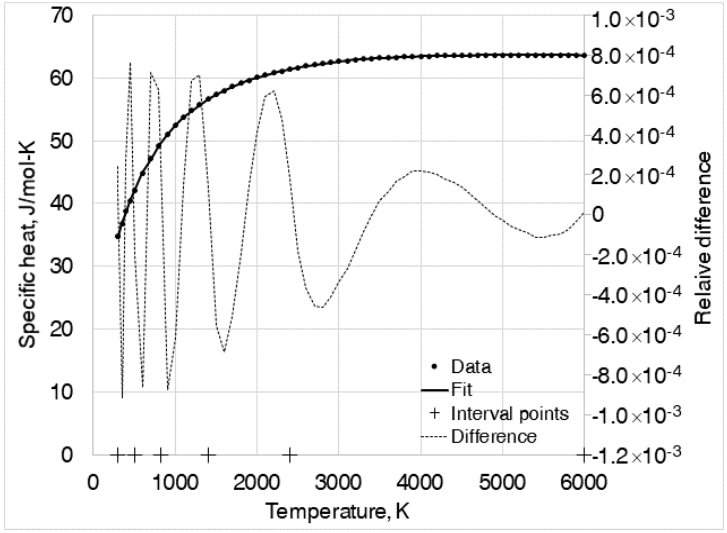


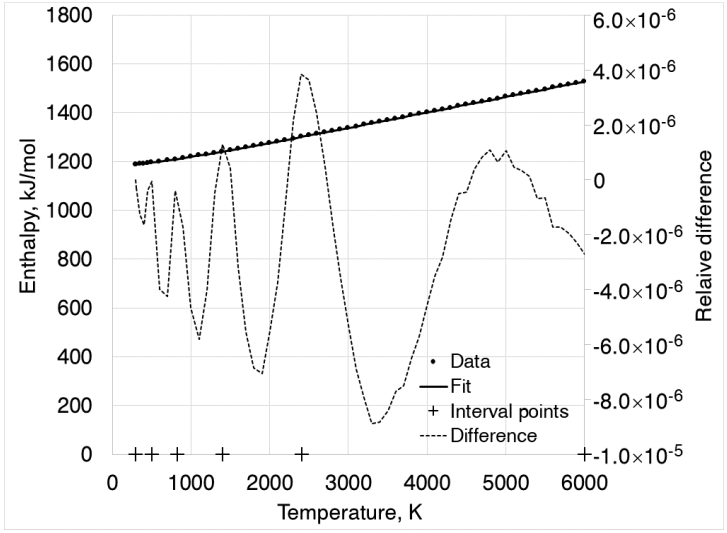


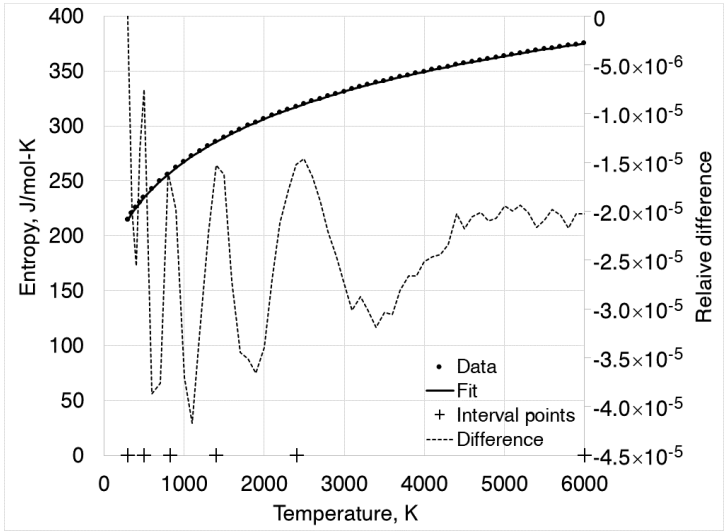


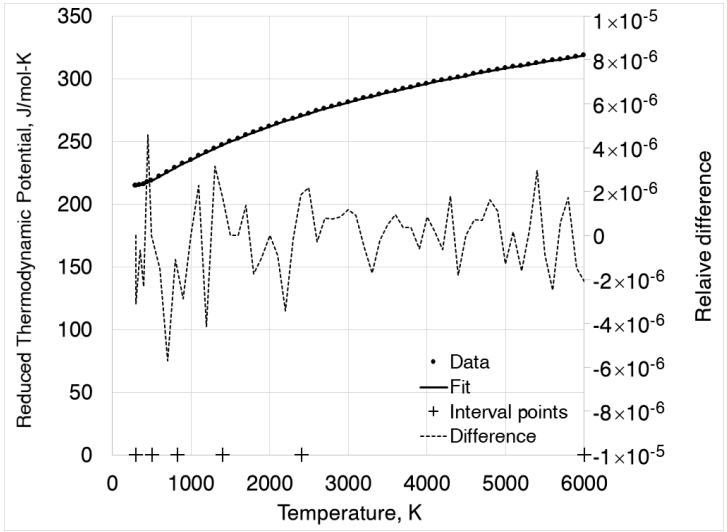
Boron Hydride Oxide, Ion (HBO^+) – B1H1O1+(g)

MW = 27.82 g/mol – $\Delta_f H^o$(298.15 K) = 1187.545 kJ/mol – S^o(298.15 K) = 214.565 J/mol-K – p^o = 0.1 MPa



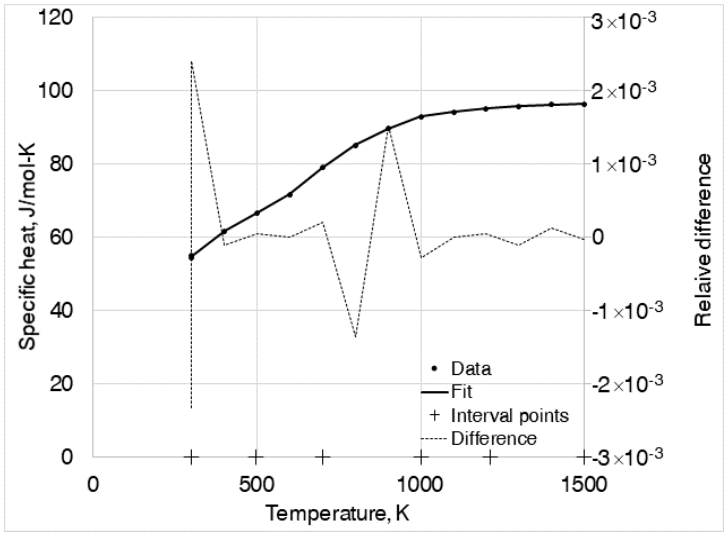


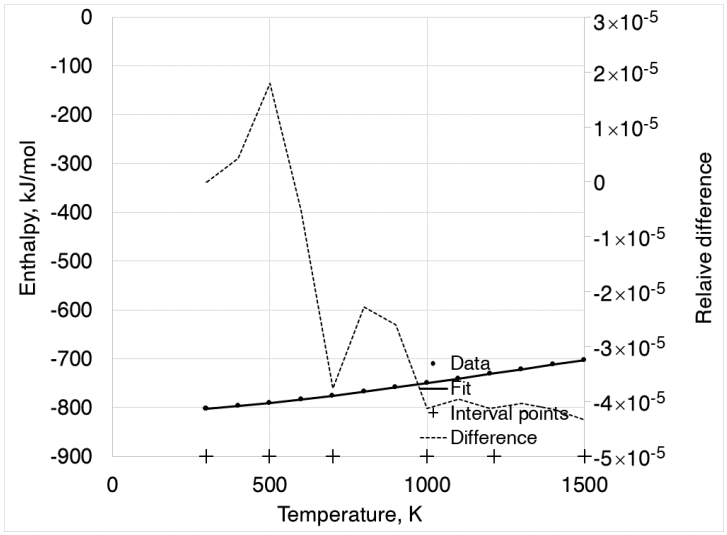


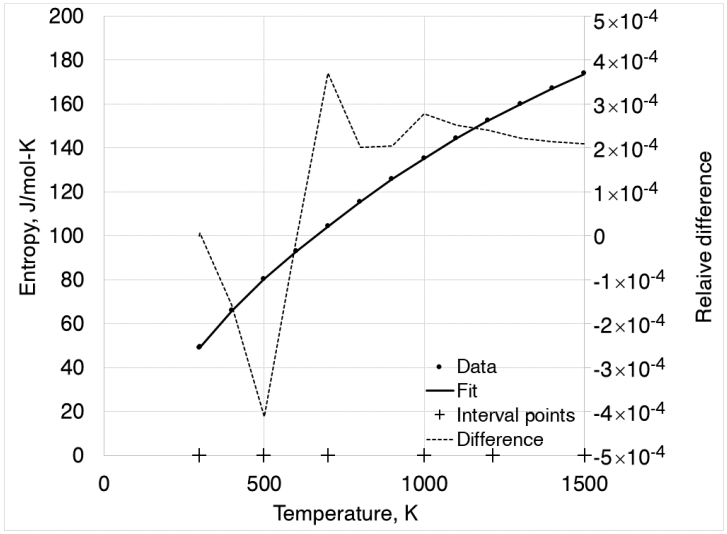


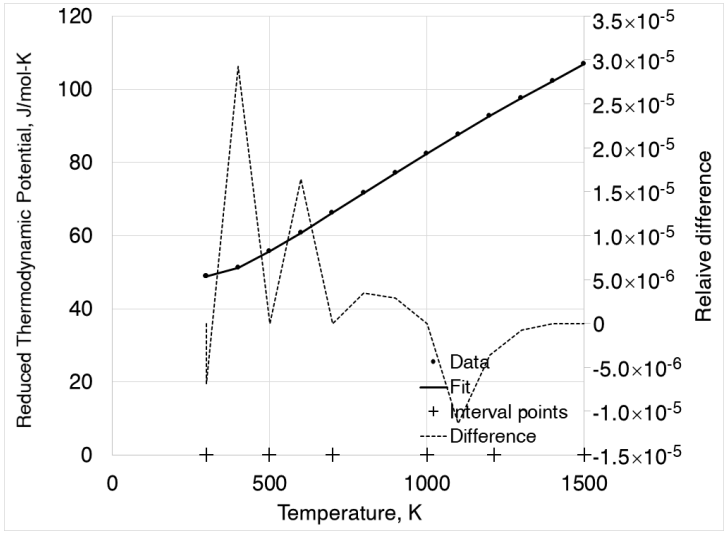
Boric Acid (HBO_2) – B1H1O2(cr)

MW = 43.82 g/mol – $\Delta_f H^o$(298.15 K) = -802.784 kJ/mol – S^o(298.15 K) = 48.953 J/mol-K – p^o = 0.1 MPa



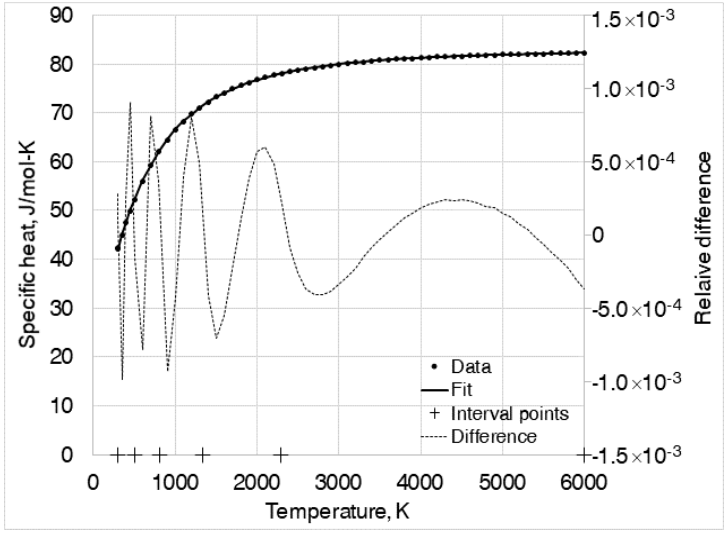


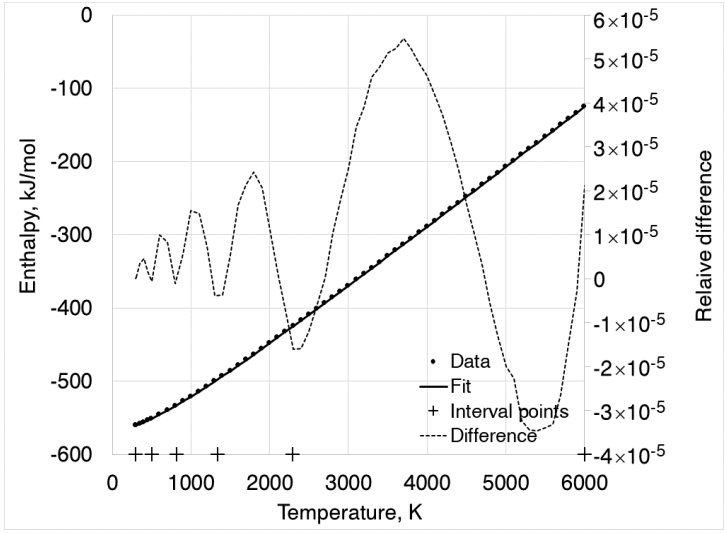


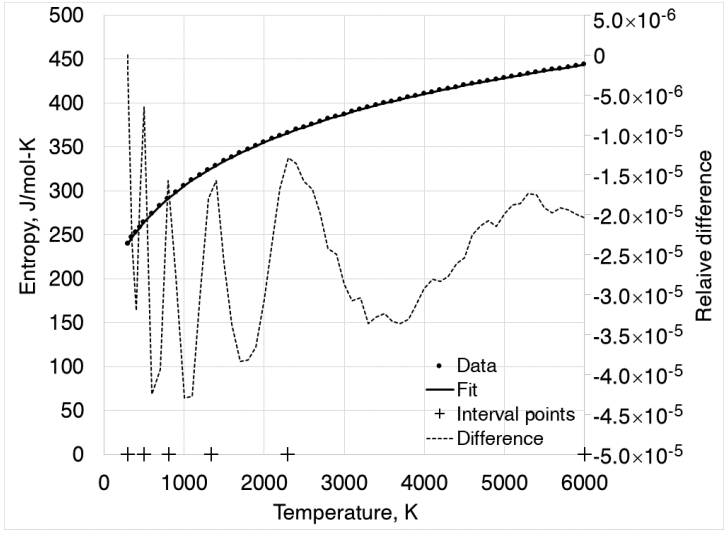


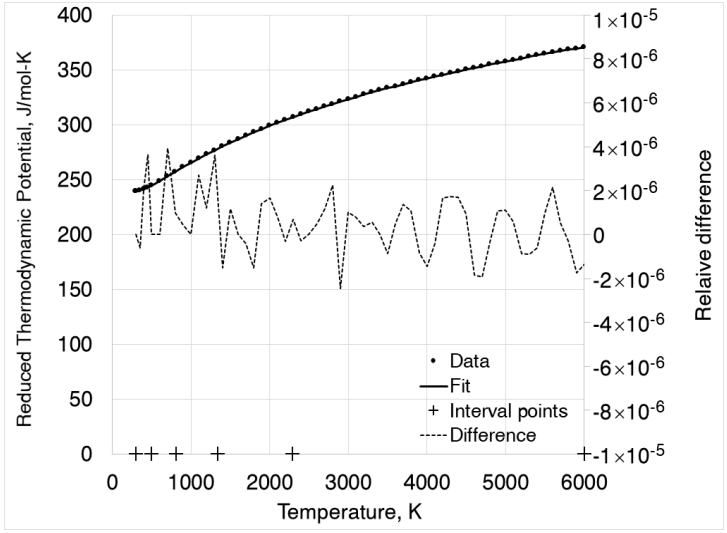
Boric Acid (HBO_2) – B1H1O2(g)

MW = 43.82 g/mol – $\Delta_f H^o$(298.15 K) = -560.656 kJ/mol – S^o(298.15 K) = 239.734 J/mol-K – p^o = 0.1 MPa



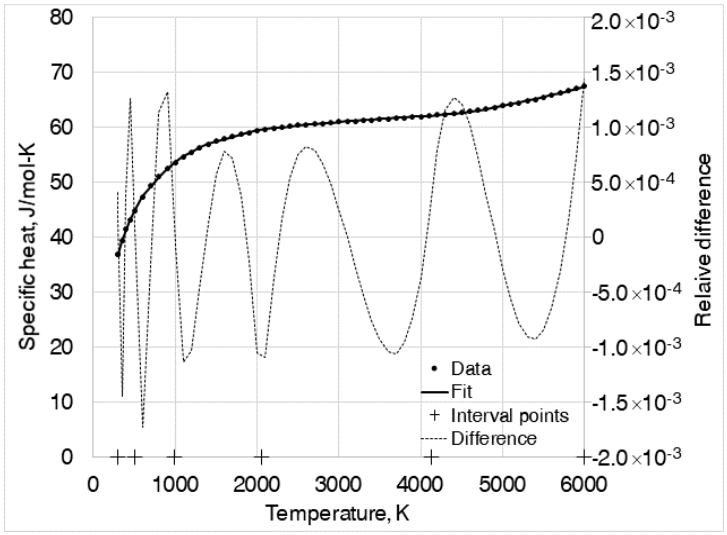


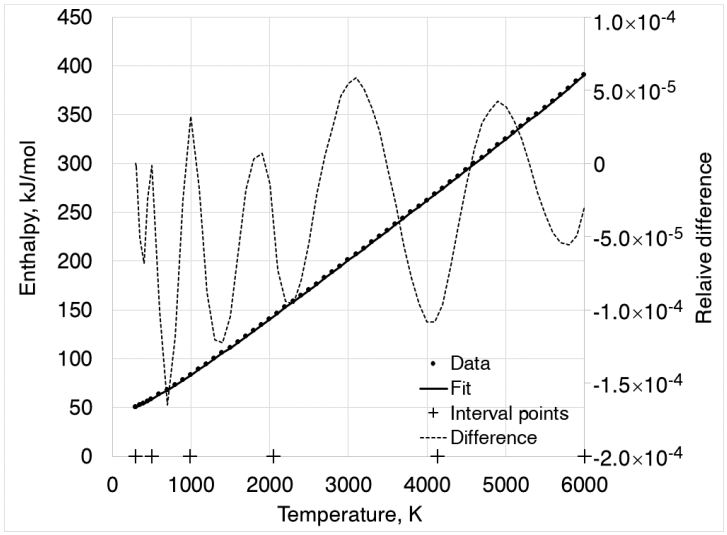


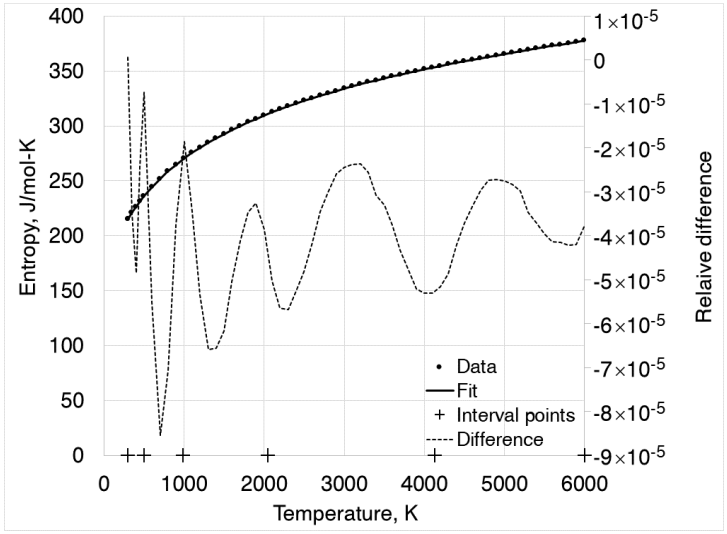


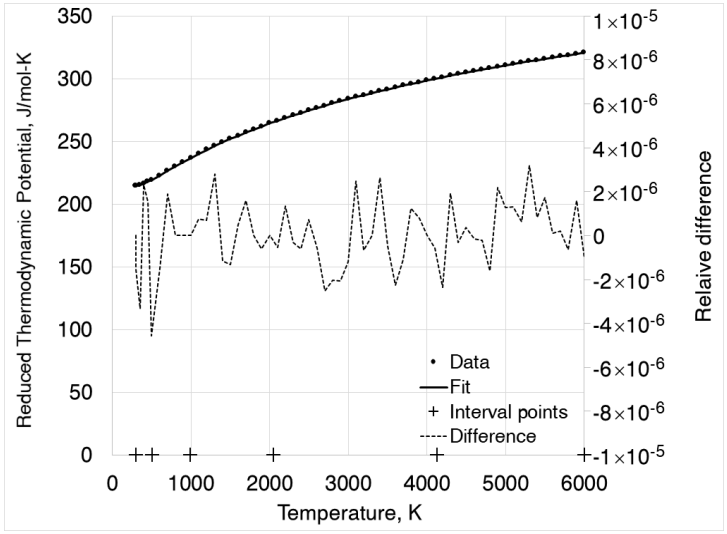
Boron Hydride Sulfide (HBS) – B1H1S1(g)

MW = 43.88 g/mol – $\Delta_f H^o$(298.15 K) = 50.208 kJ/mol – S^o(298.15 K) = 214.964 J/mol-K – p^o = 0.1 MPa



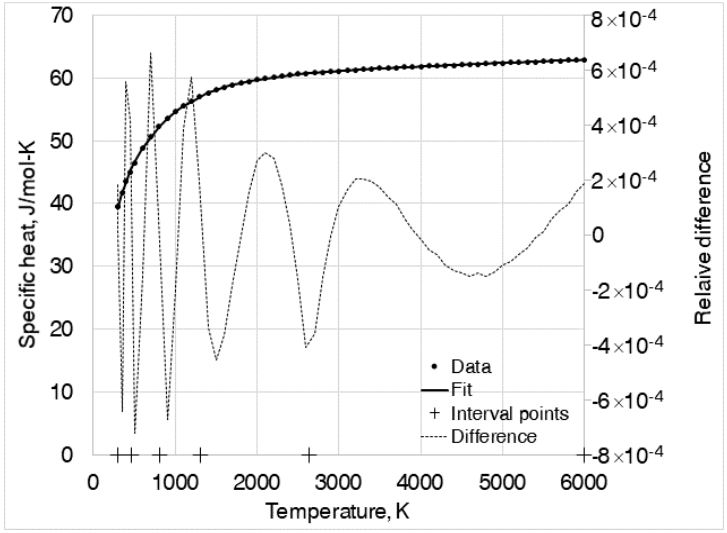


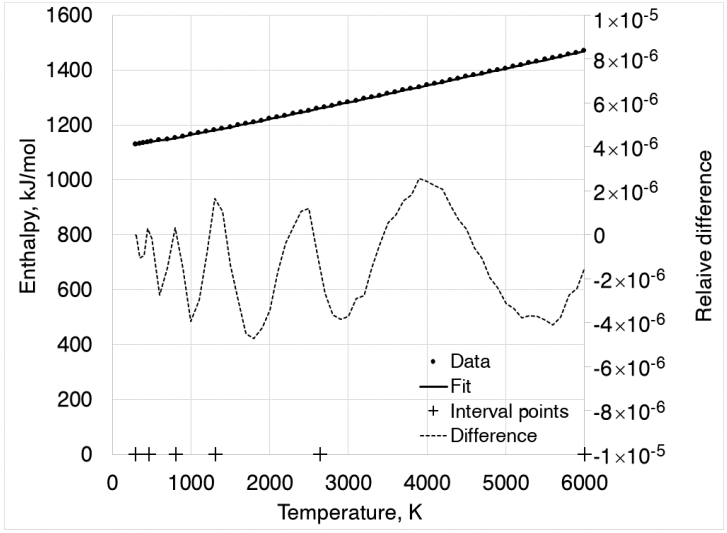


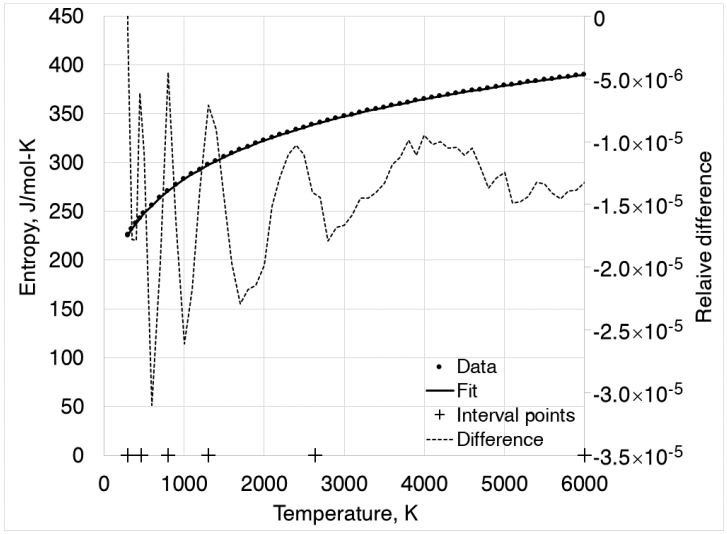


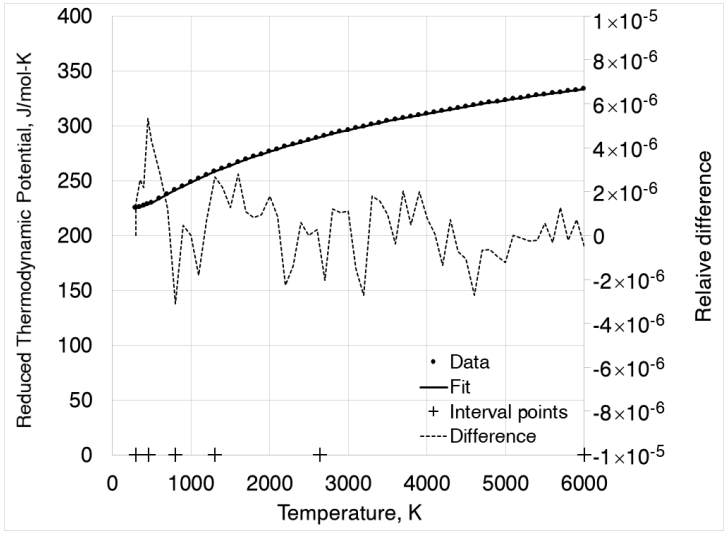
Boron Hydride Sulfide, Ion (HBS^+) – B1H1S1+(g)

MW = 43.88 g/mol – $\Delta_f H^o$(298.15 K) = 1129.471 kJ/mol – S^o(298.15 K) = 225.35 J/mol-K – p^o = 0.1 MPa



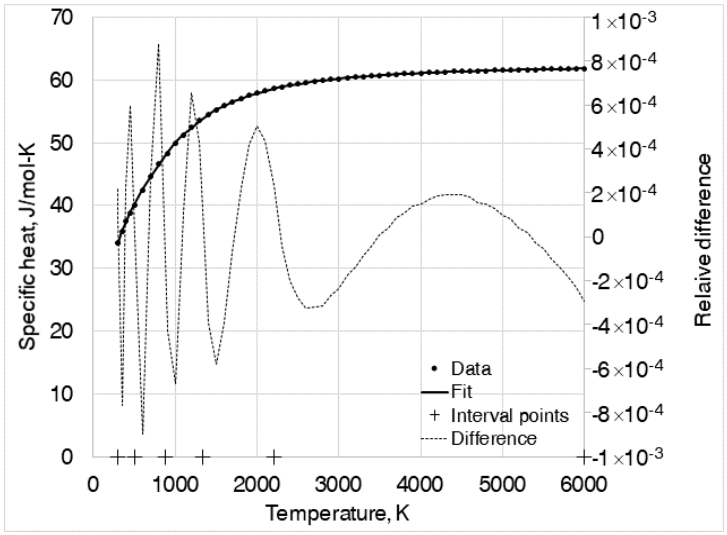


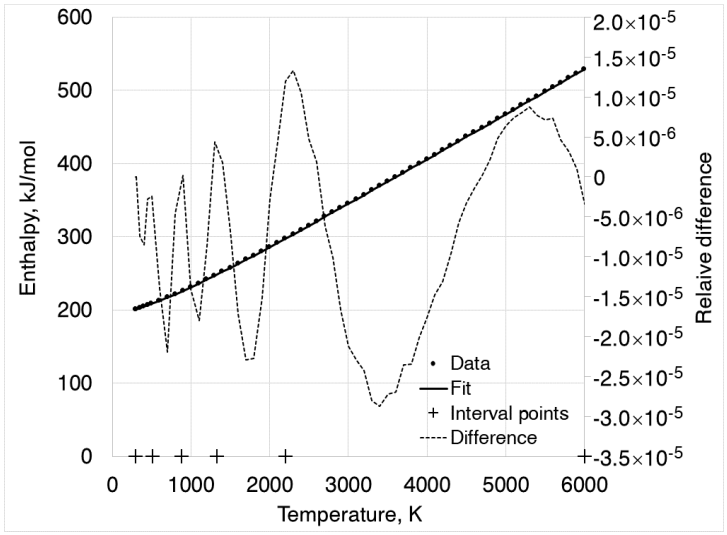


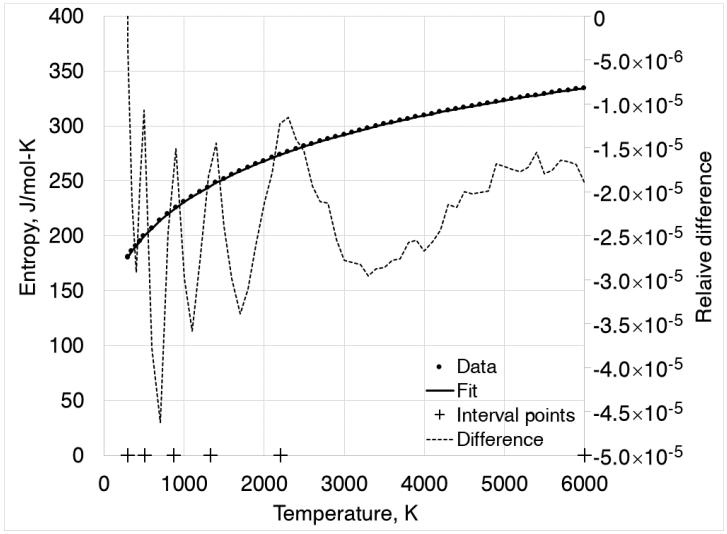


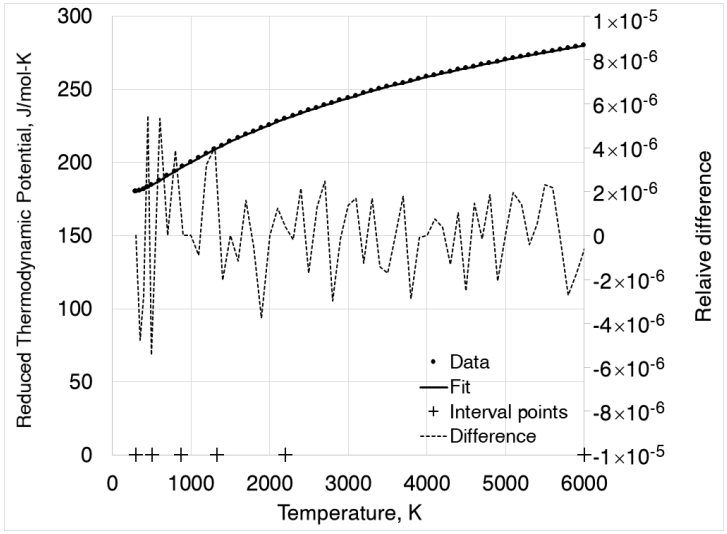
Borane (BH_2) – B1H2(g)

MW = 12.83 g/mol – $\Delta_f H^o$(298.15 K) = 200.832 kJ/mol – S^o(298.15 K) = 180.195 J/mol-K – p^o = 0.1 MPa



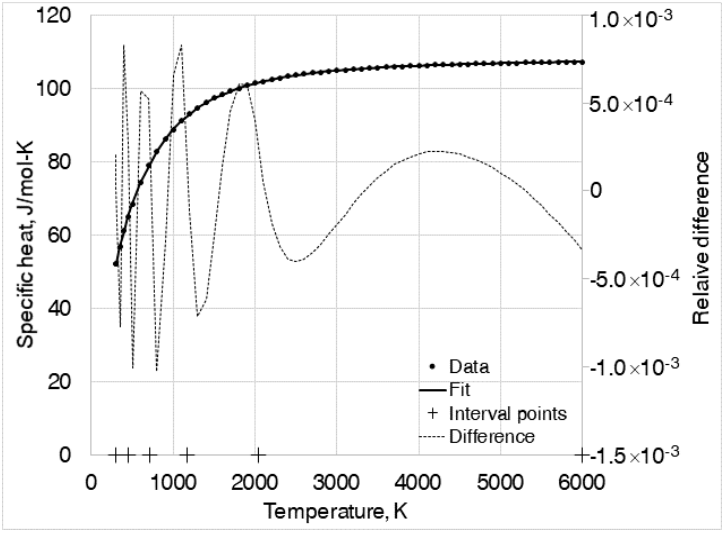


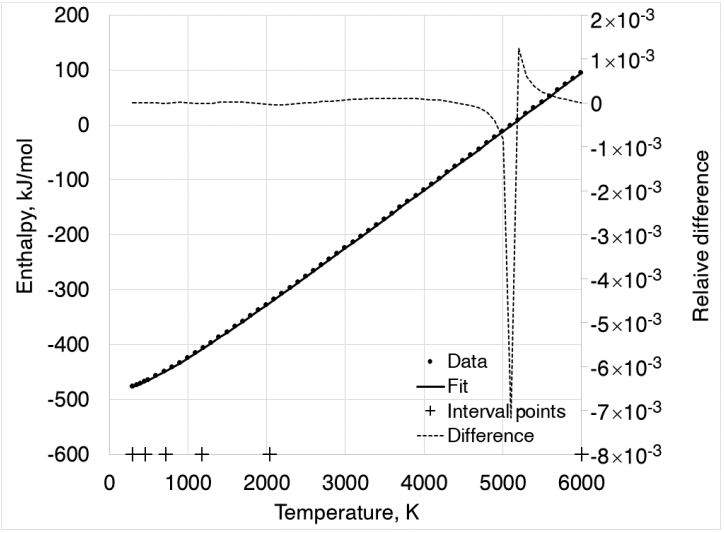


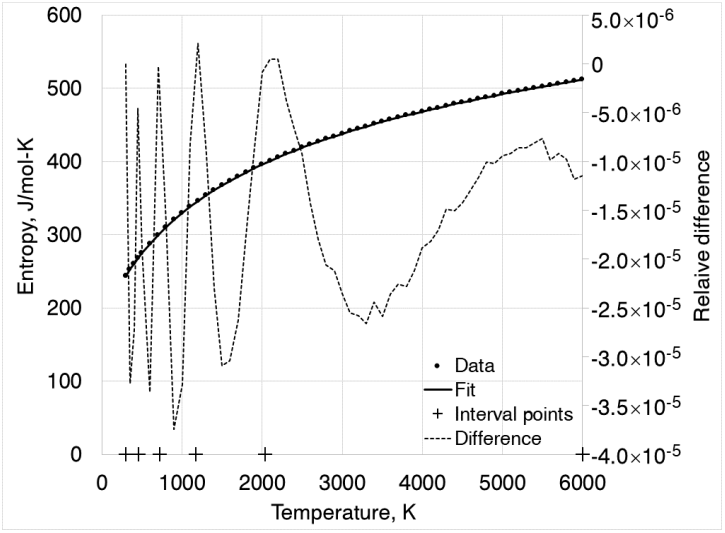


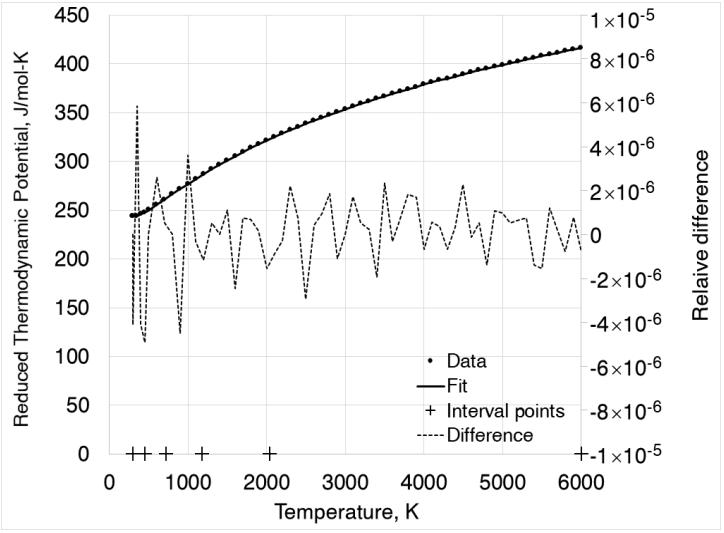
Dihydroxyborane ($B(OH)_2$) – B1H2O2(g)

MW = 44.82 g/mol – $\Delta_f H^o$(298.15 K) = -476.976 kJ/mol – S^o(298.15 K) = 243.589 J/mol-K – p^o = 0.1 MPa



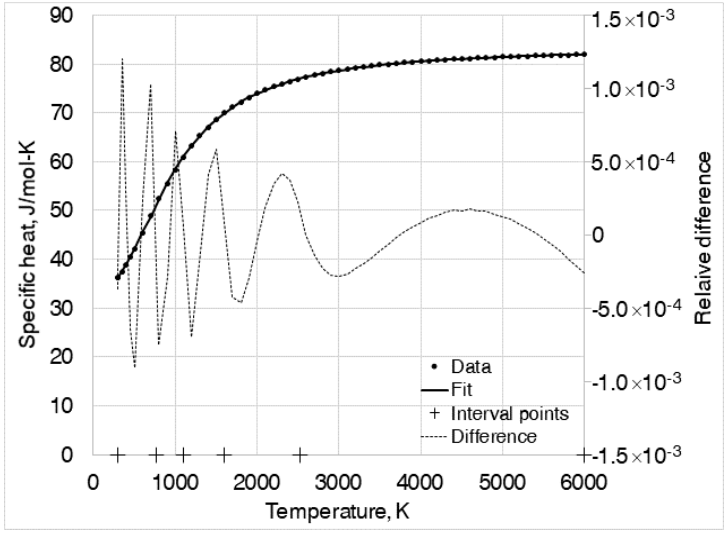


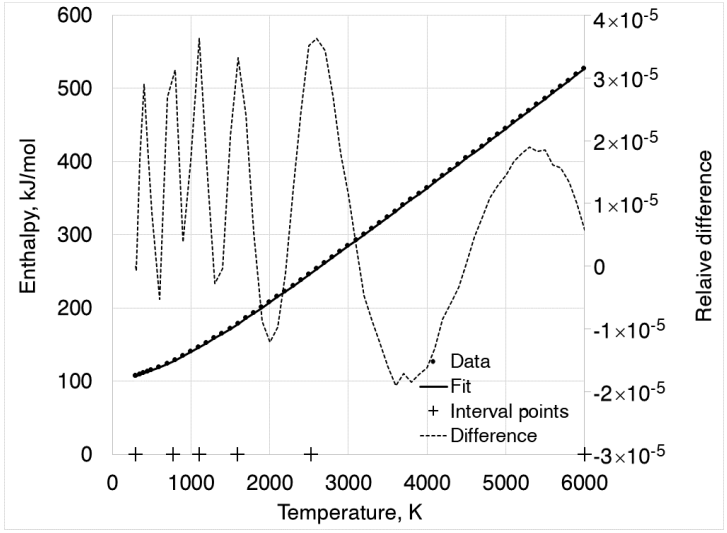


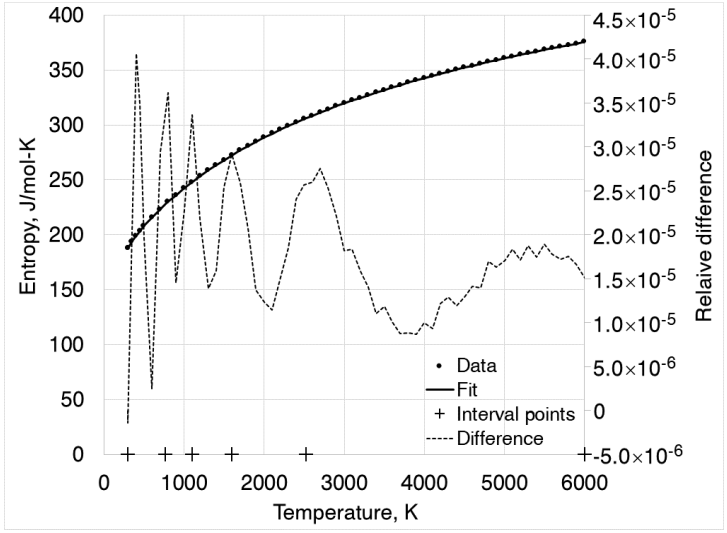


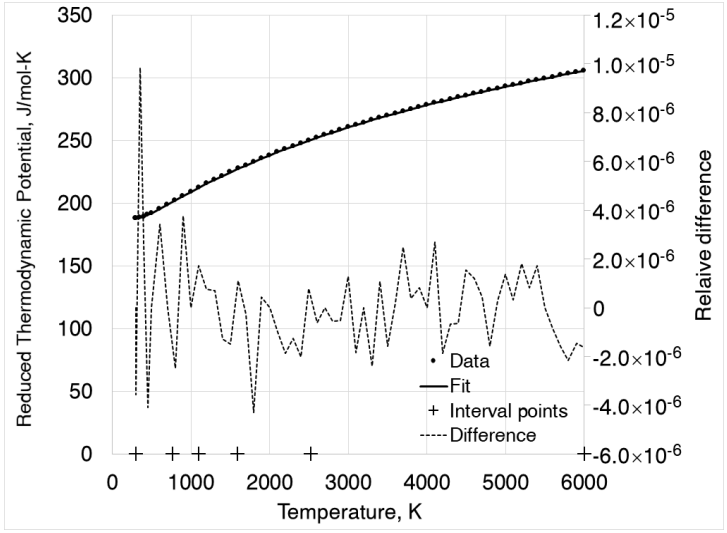
Borane (BH_3) – B1H3(g)

MW = 13.83 g/mol – $\Delta_f H^o$(298.15 K) = 106.692 kJ/mol – S^o(298.15 K) = 187.876 J/mol-K – p^o = 0.1 MPa



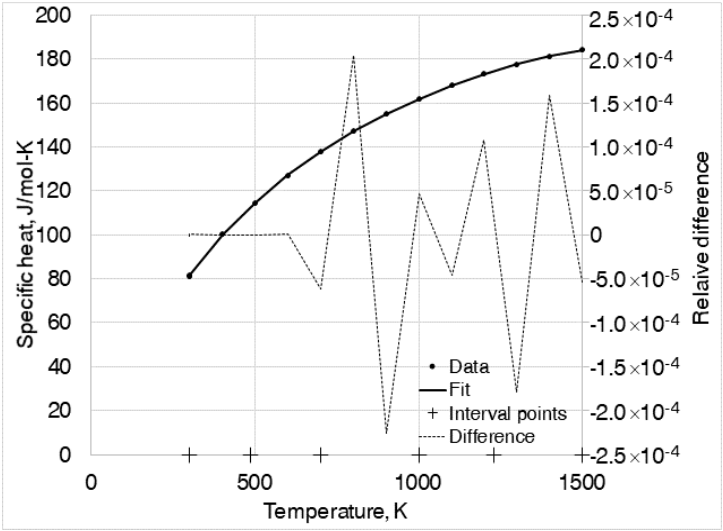


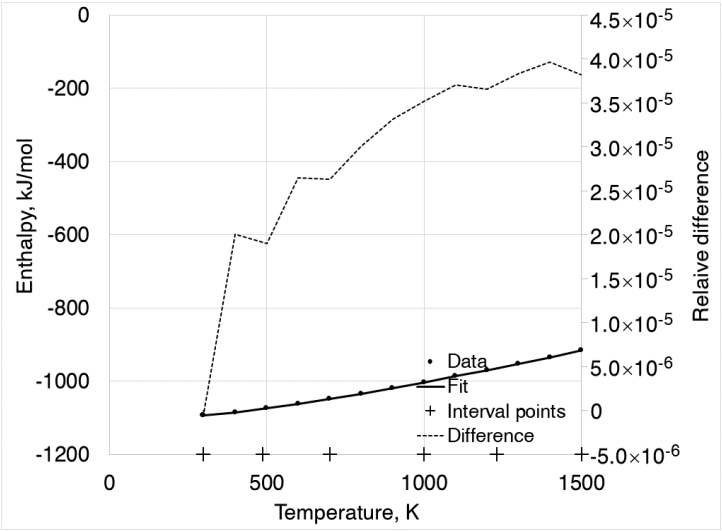


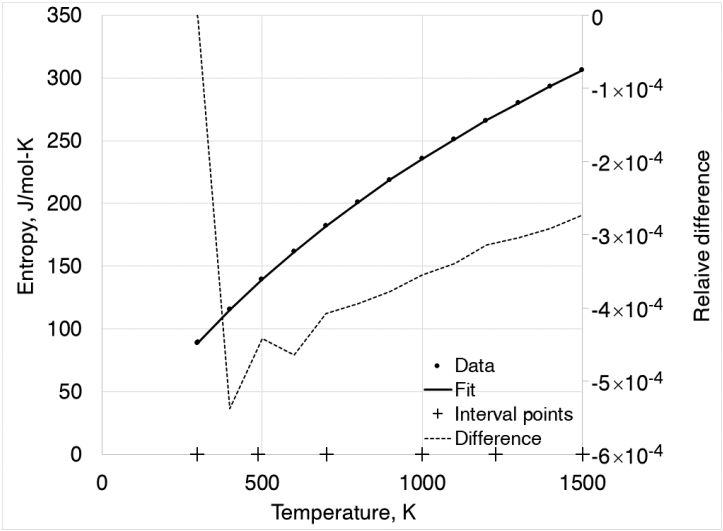


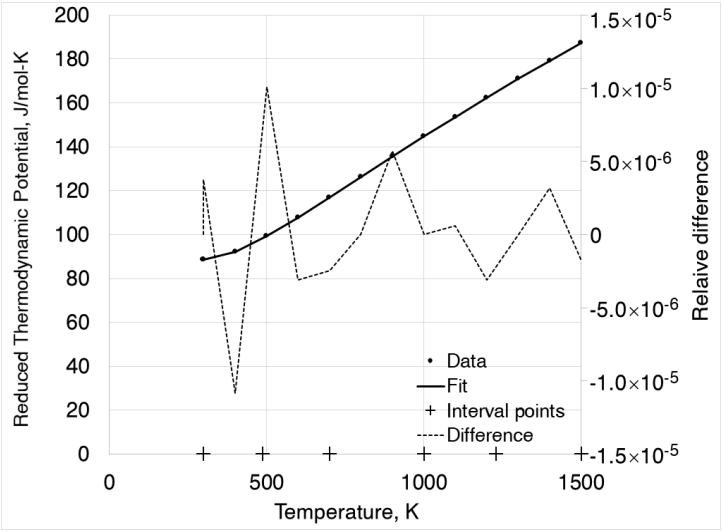
Boric Acid (H_3BO_3) – B1H3O3(cr)

MW = 61.83 g/mol – $\Delta_f H^o$(298.15 K) = -1093.99 kJ/mol – S^o(298.15 K) = 88.743 J/mol-K – p^o = 0.1 MPa



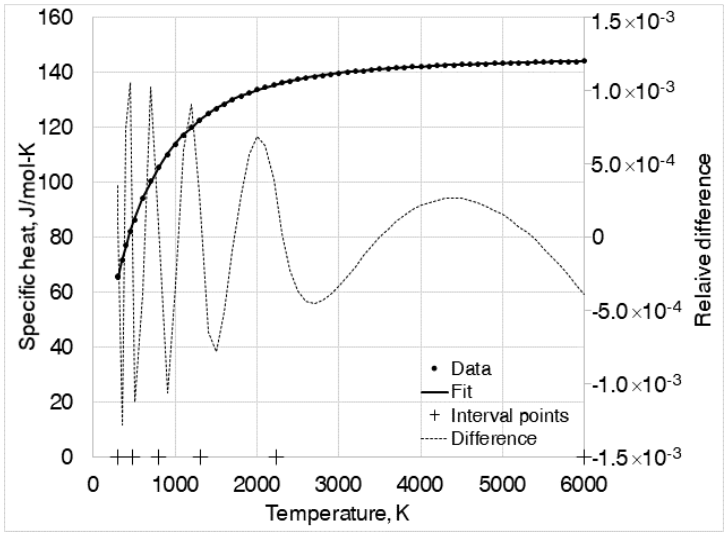


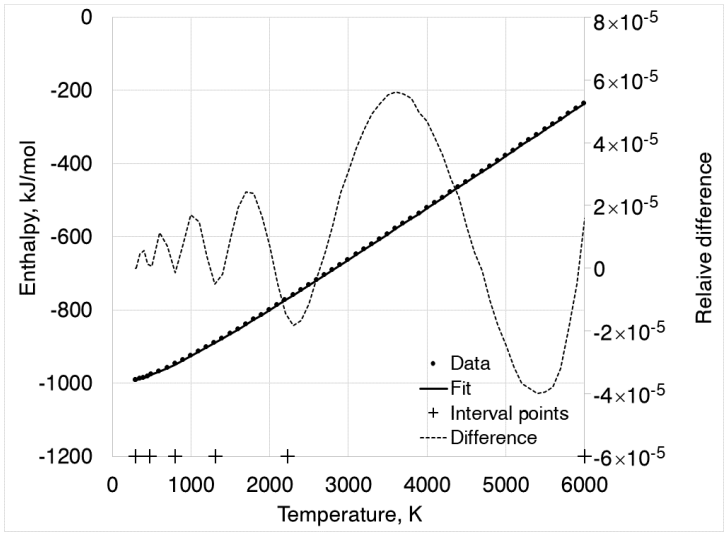


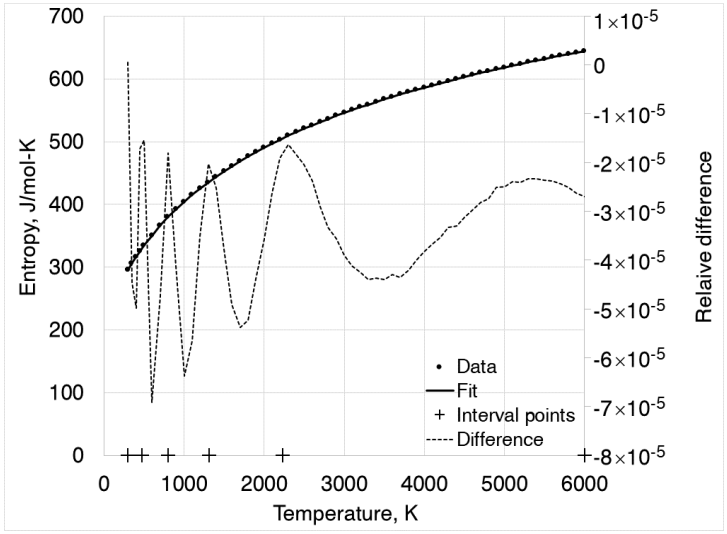


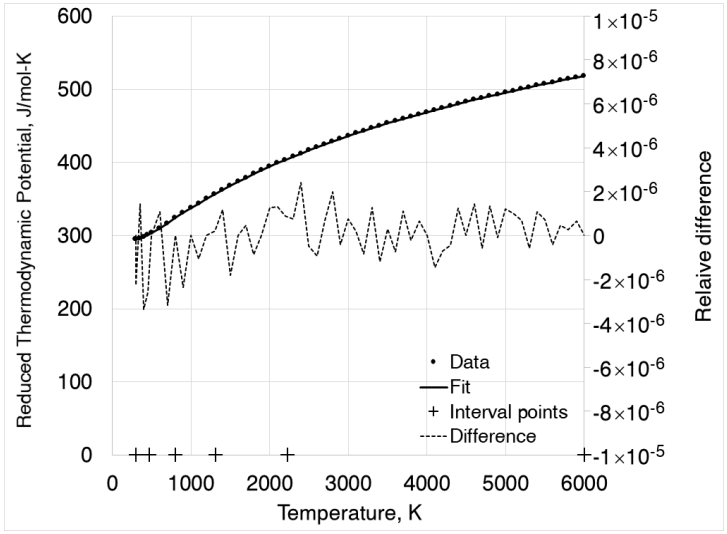
Boric Acid (H_3BO_3) – B1H3O3(g)

MW = 61.83 g/mol – $\Delta_f H^o$(298.15 K) = -992.277 kJ/mol – S^o(298.15 K) = 295.237 J/mol-K – p^o = 0.1 MPa



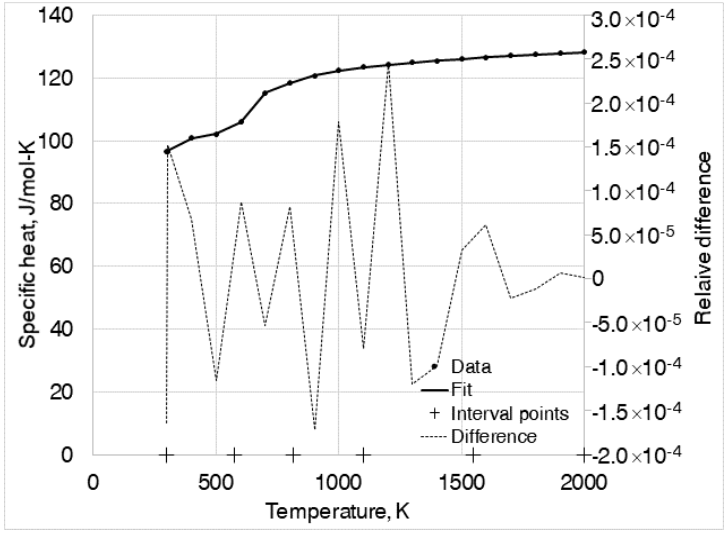


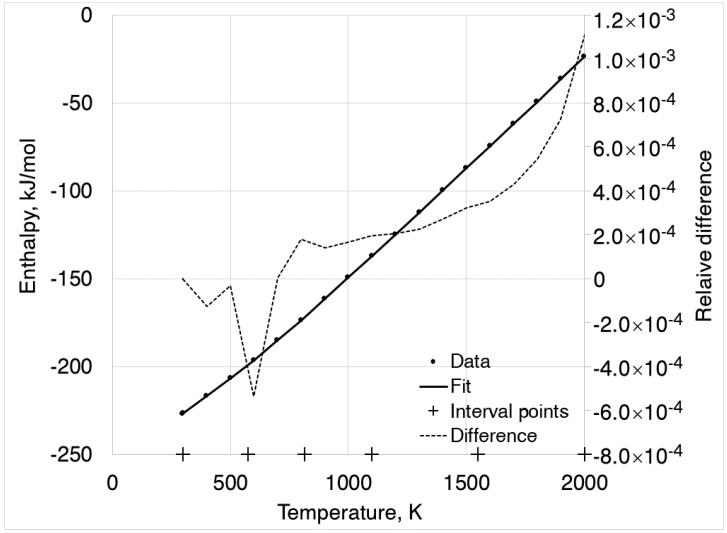


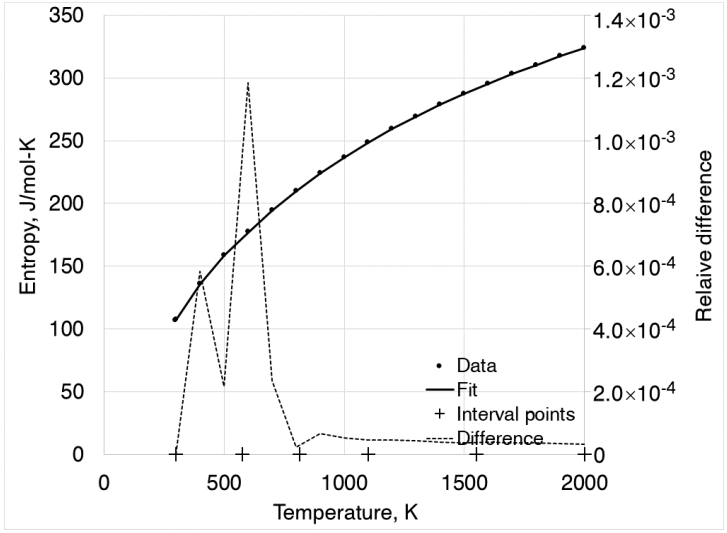


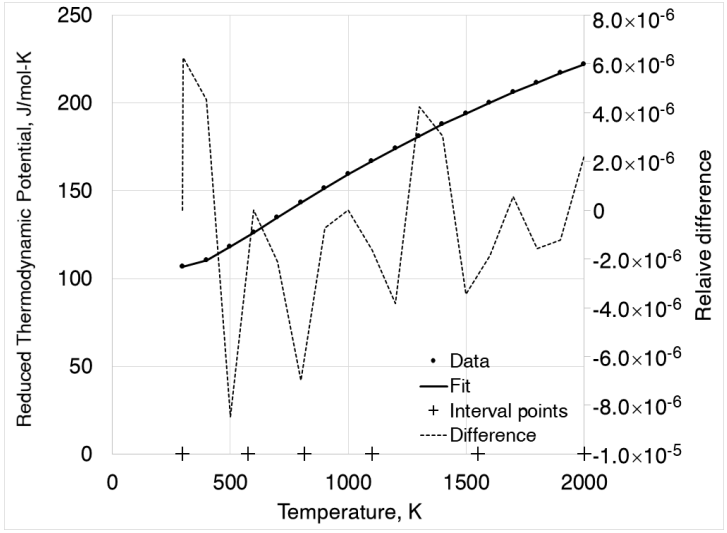
Potassium Tetrahydroborate (KBH_4) – B1H4K1(cr)

MW = 53.94 g/mol – $\Delta_f H^o$(298.15 K) = -226.898 kJ/mol – S^o(298.15 K) = 106.608 J/mol-K – p^o = 0.1 MPa



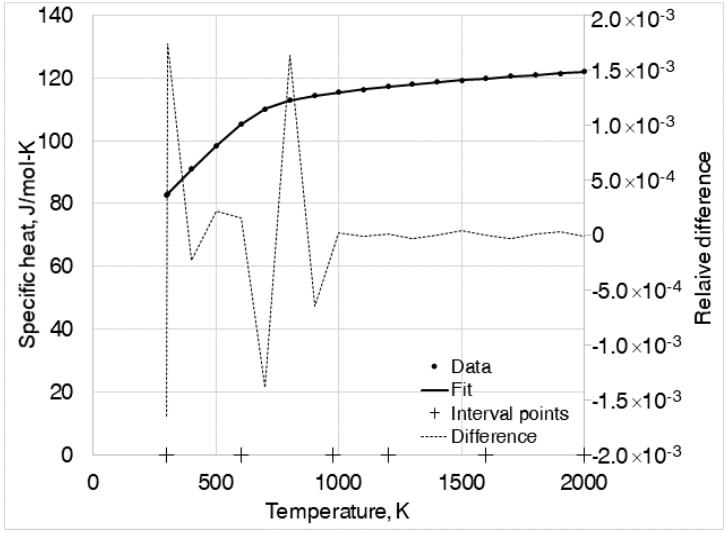


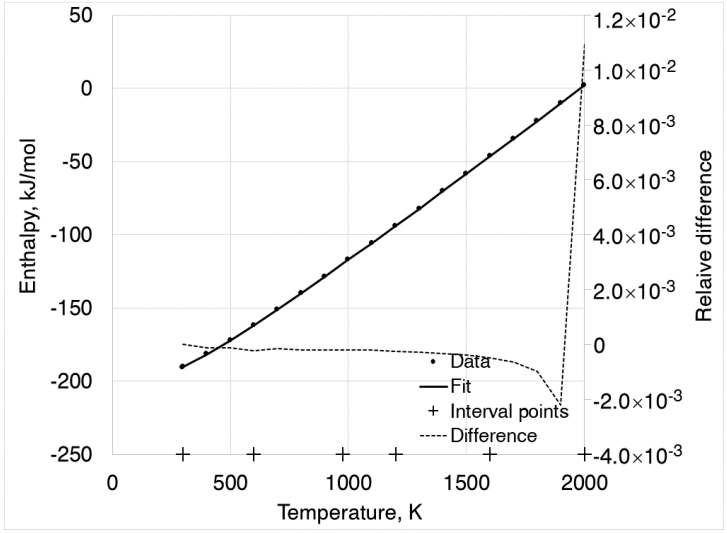


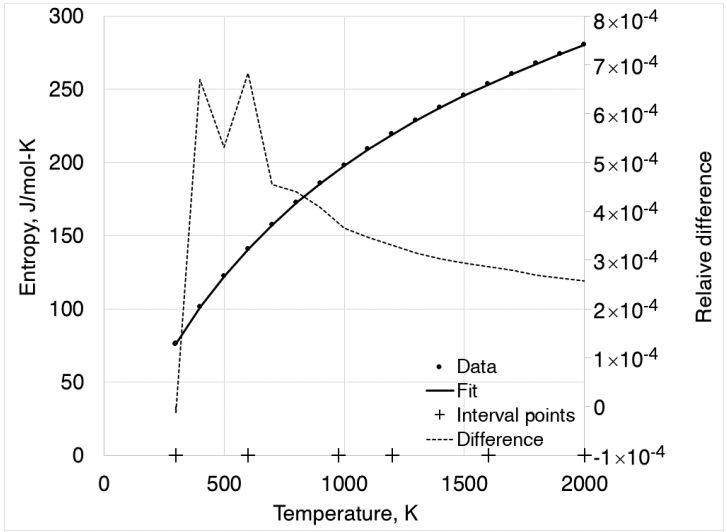


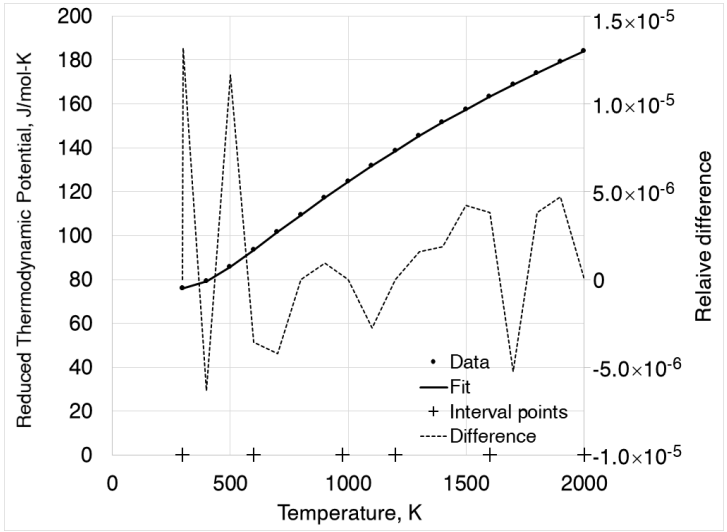
Lithium Tetrahydroborate ($LiBH_4$) – B1H4Li1(cr)

MW = 21.78 g/mol – $\Delta_f H^o$(298.15 K) = -190.464 kJ/mol – S^o(298.15 K) = 75.815 J/mol-K – p^o = 0.1 MPa



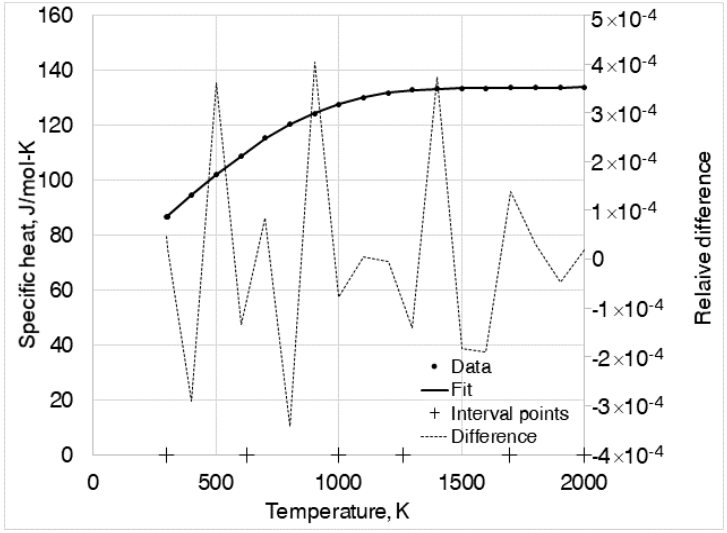


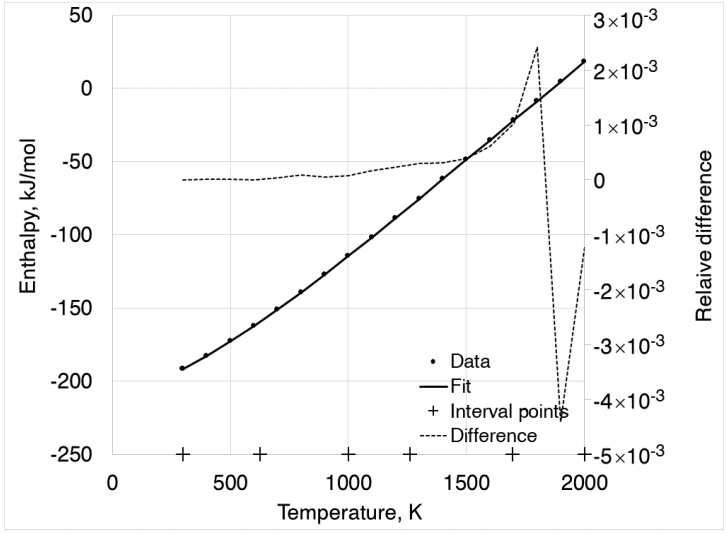


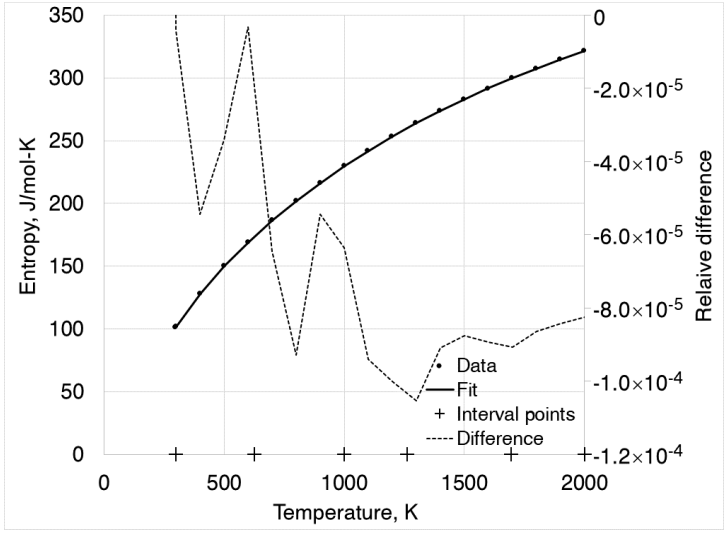


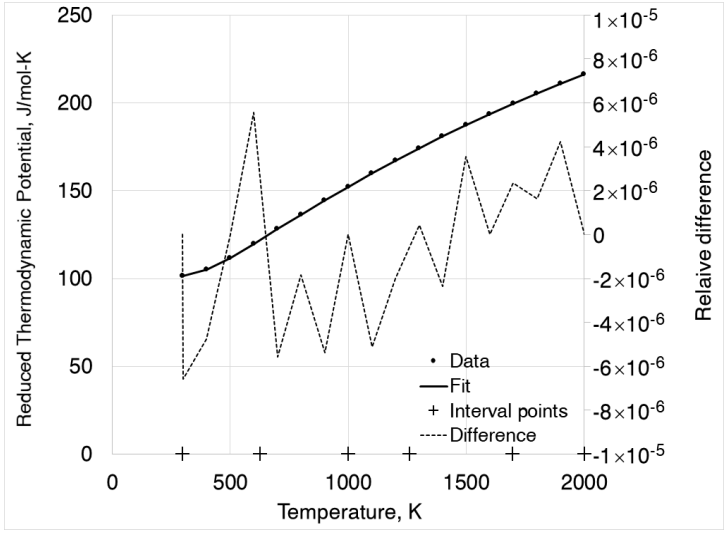
Sodium Tetrahydroborate ($NaBH_4$) – B1H4Na1(cr)

MW = 37.83 g/mol – $\Delta_f H^o$(298.15 K) = -191.836 kJ/mol – S^o(298.15 K) = 101.387 J/mol-K – p^o = 0.1 MPa



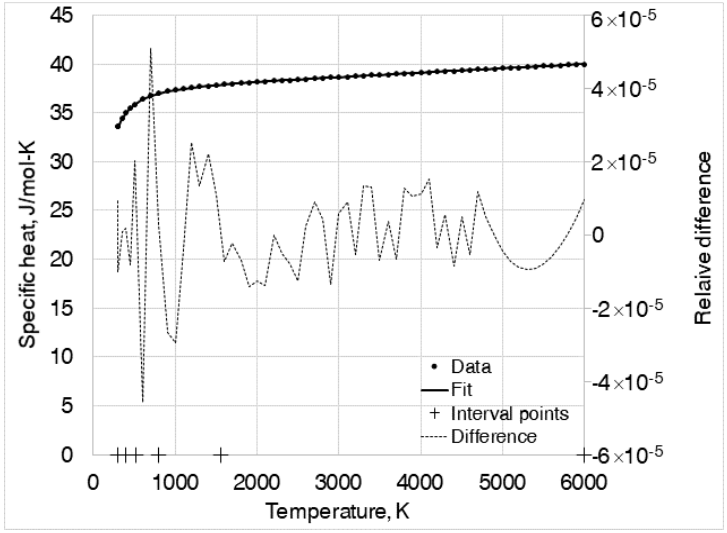


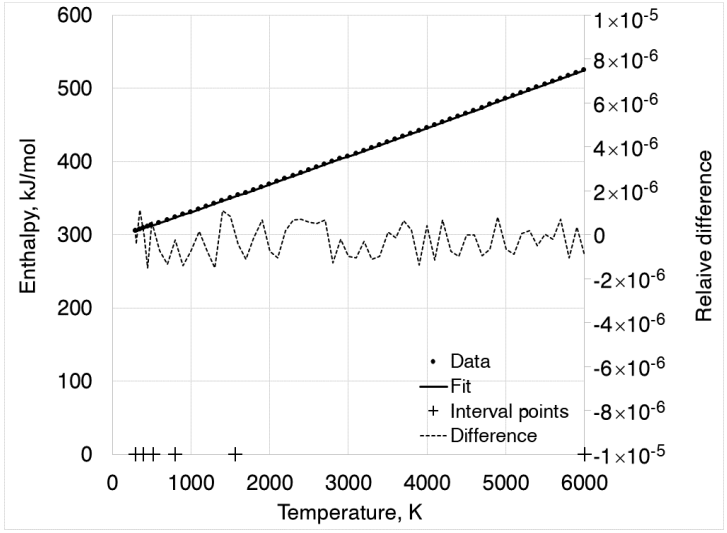


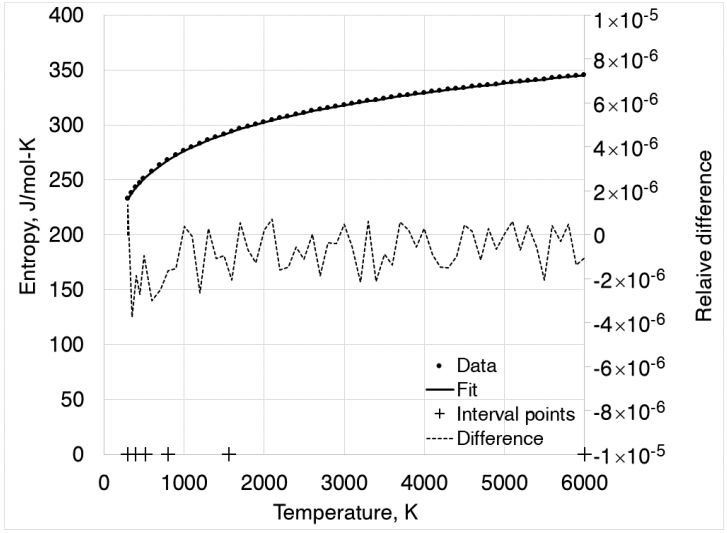


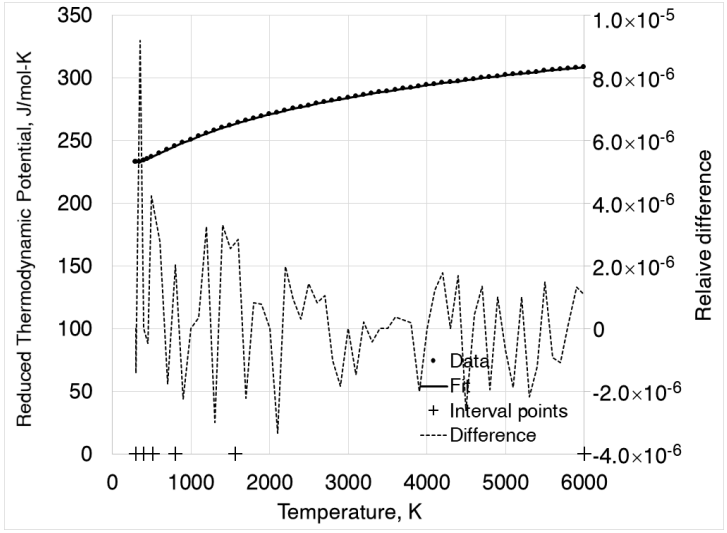
Iodoborane (BI) – B1I1(g)

MW = 137.71 g/mol – $\Delta_f H^o$(298.15 K) = 305.432 kJ/mol – S^o(298.15 K) = 232.75 J/mol-K – p^o = 0.1 MPa



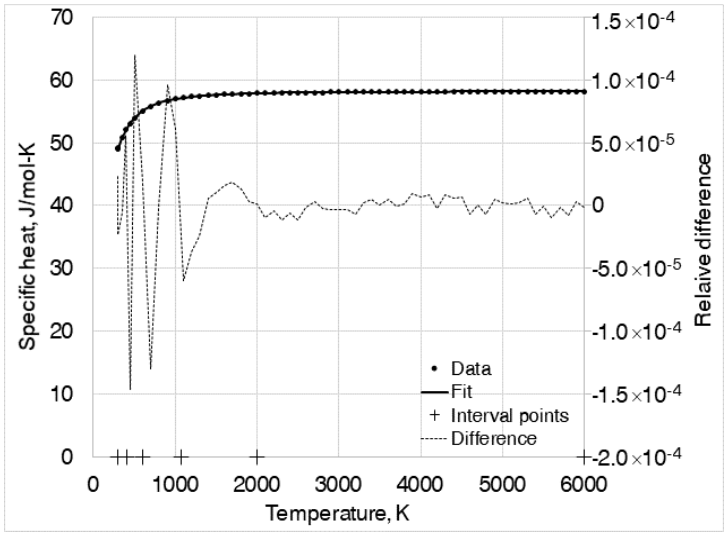


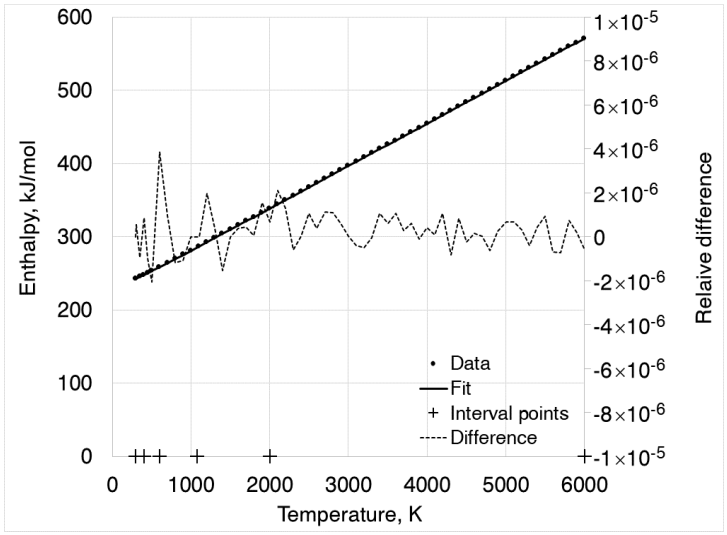


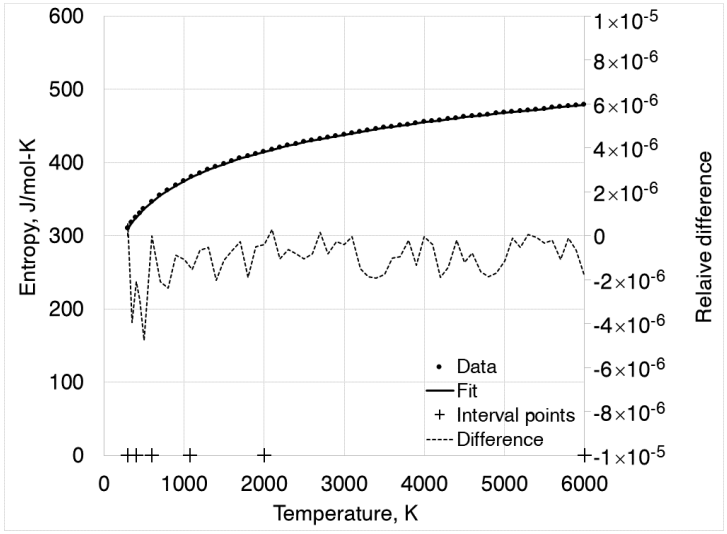


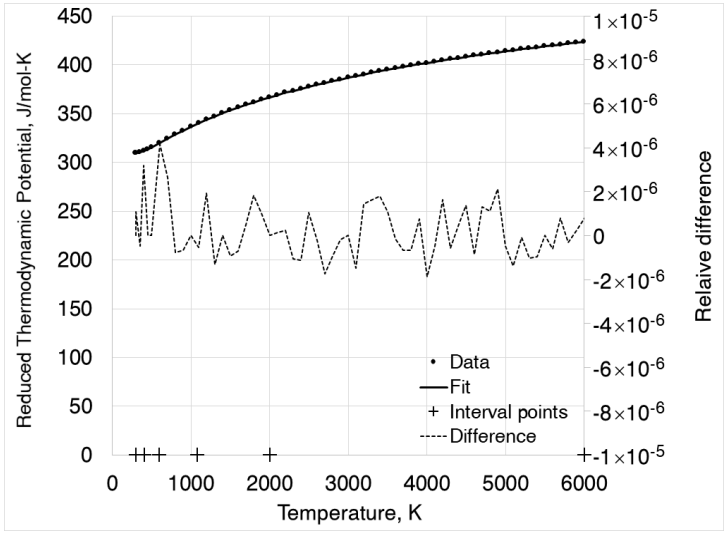
Diiodoborane (BI_2) – B1I2(g)

MW = 264.62 g/mol – $\Delta_f H^o$(298.15 K) = 242.672 kJ/mol – S^o(298.15 K) = 309.56 J/mol-K – p^o = 0.1 MPa



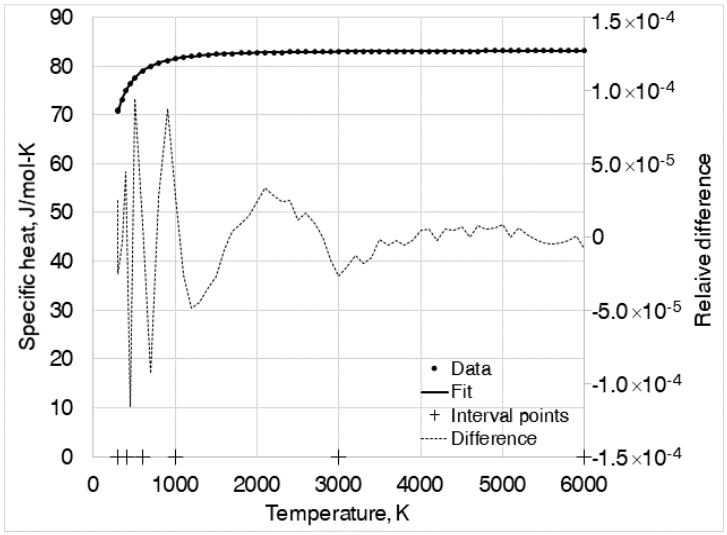


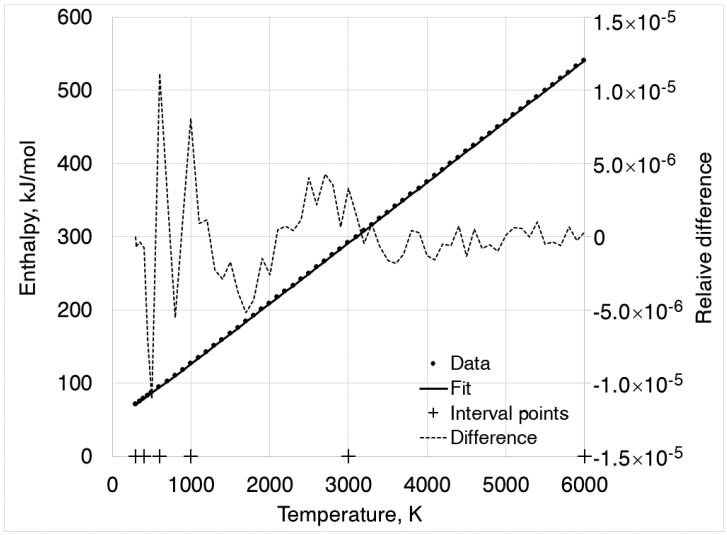


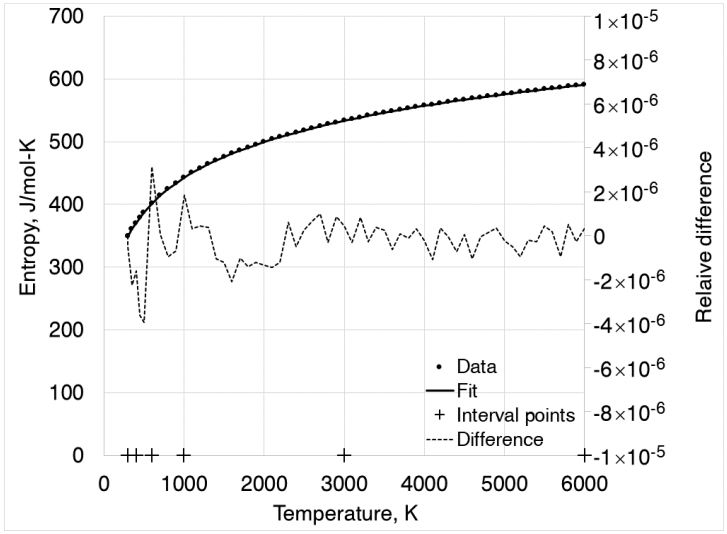


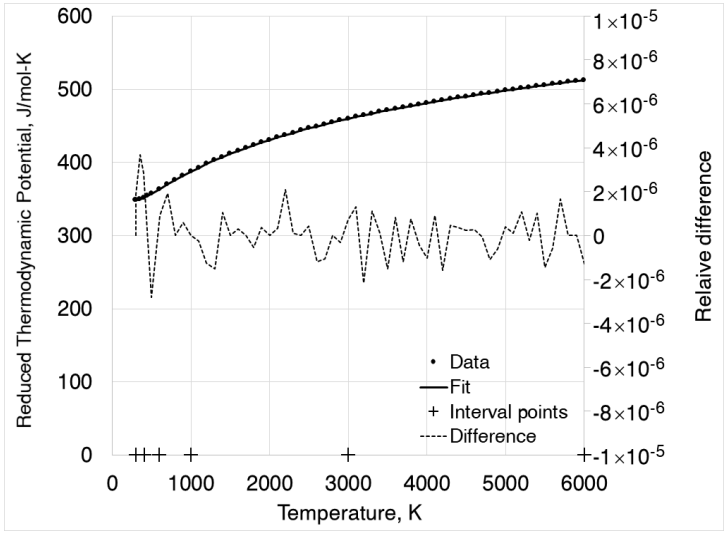
Triiodoborane (BI_3) – B1I3(g)

MW = 391.52 g/mol – $\Delta_f H^o$(298.15 K) = 71.128 kJ/mol – S^o(298.15 K) = 348.721 J/mol-K – p^o = 0.1 MPa



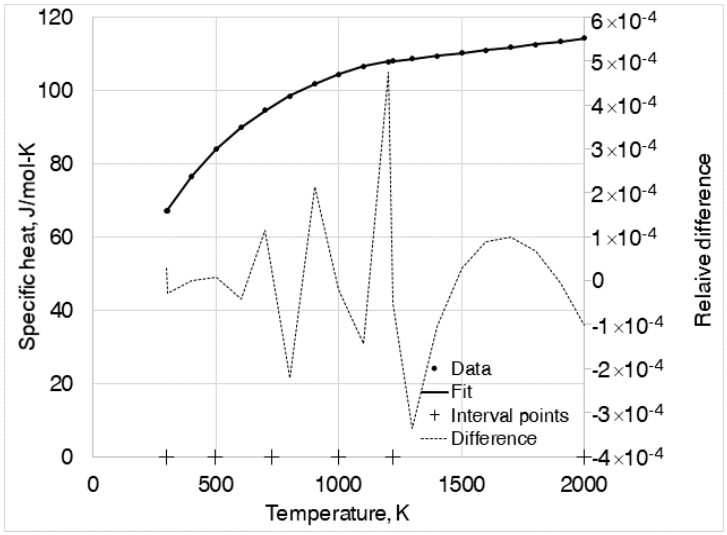


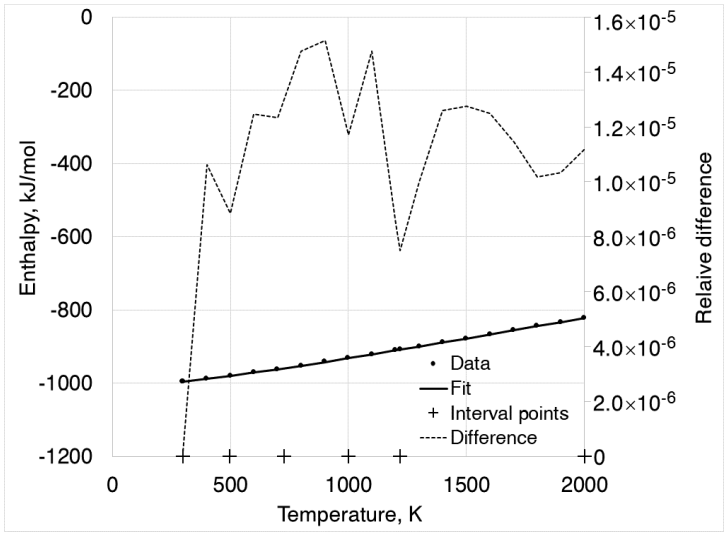


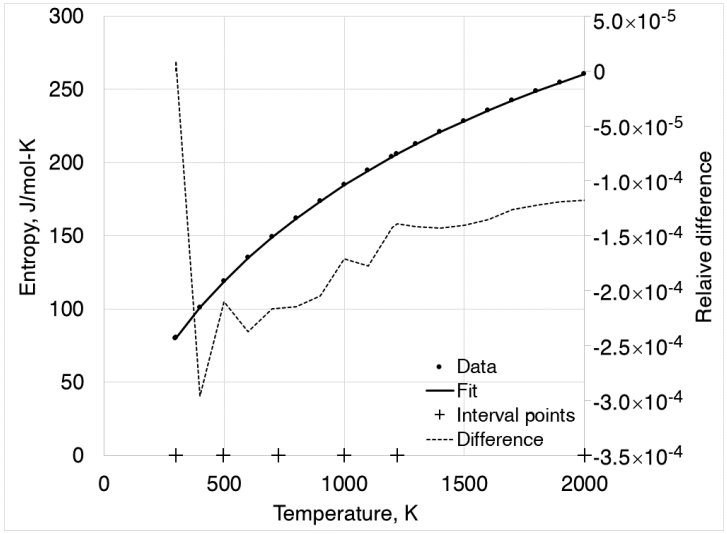


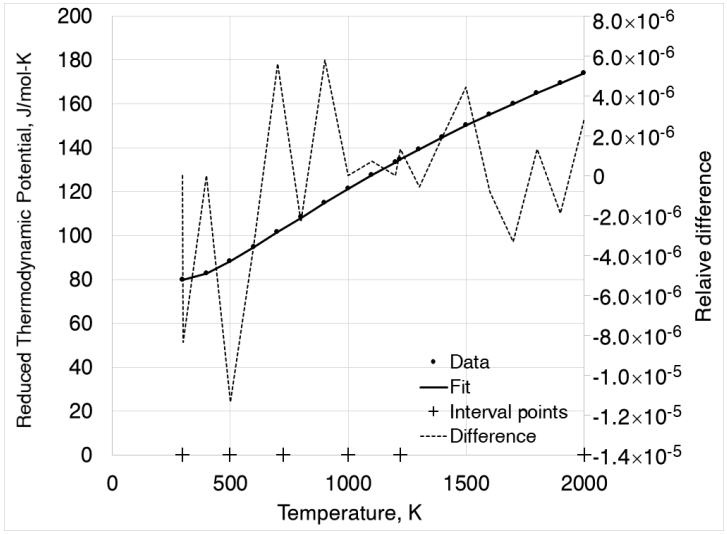
Potassium Borate (KBO_2) – B1K1O2(cr)

MW = 81.91 g/mol – $\Delta_f H^o$(298.15 K) = -994.955 kJ/mol – S^o(298.15 K) = 79.981 J/mol-K – p^o = 0.1 MPa



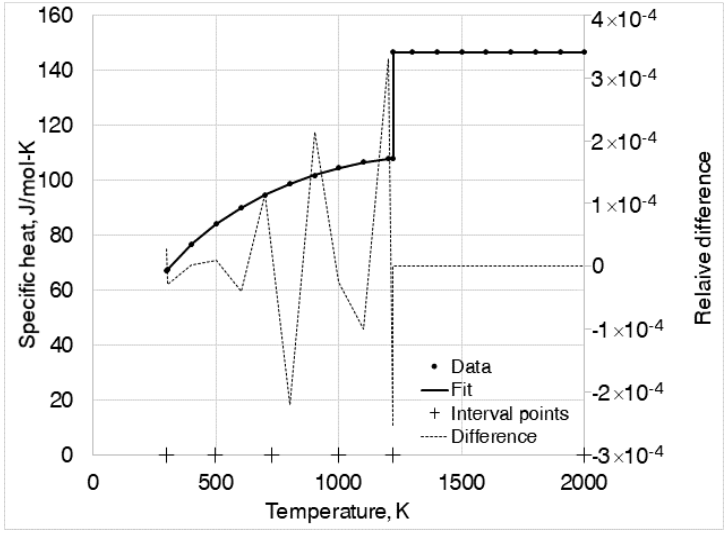


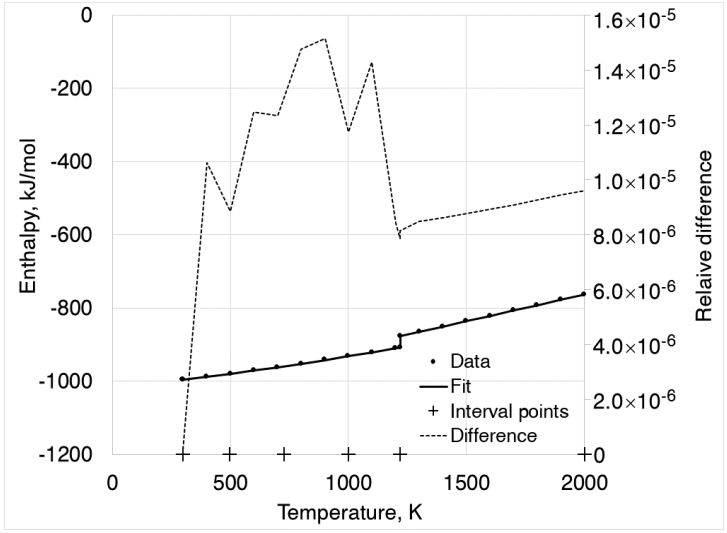


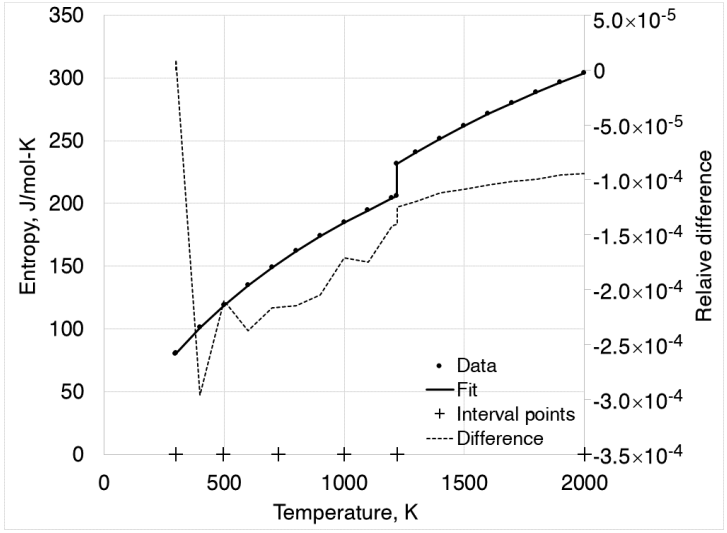


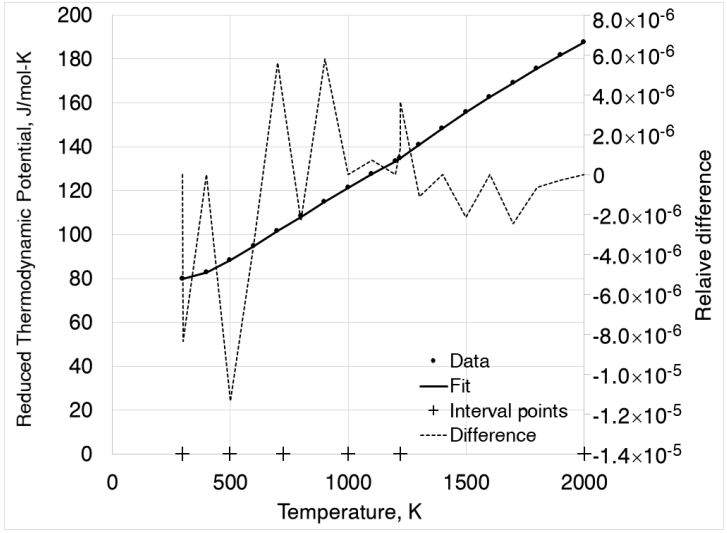
Potassium Borate (KBO_2) – B1K1O2(cr,l)

MW = 81.91 g/mol – $\Delta_f H^o$(298.15 K) = -994.955 kJ/mol – S^o(298.15 K) = 79.981 J/mol-K – p^o = 0.1 MPa



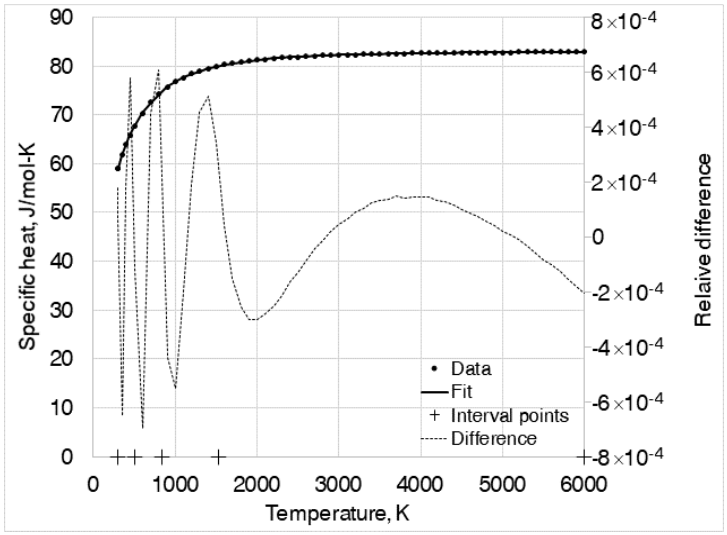


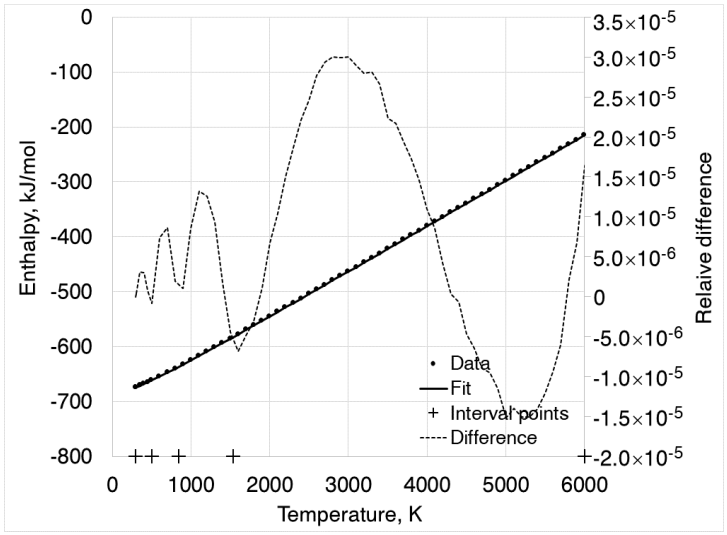


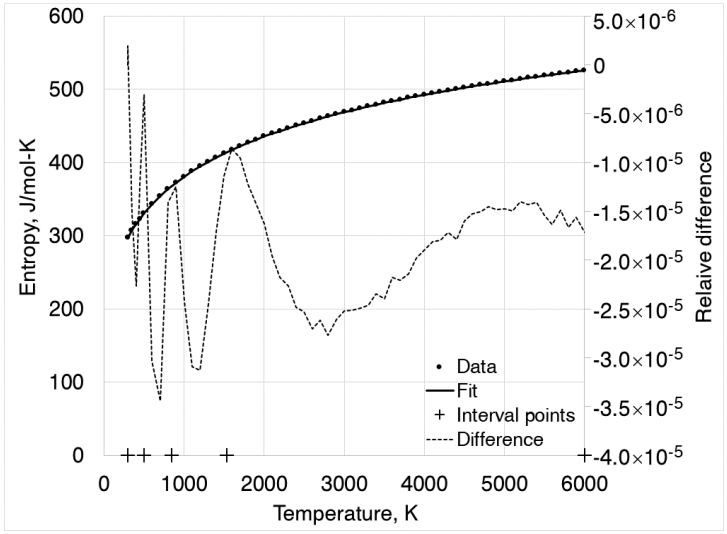


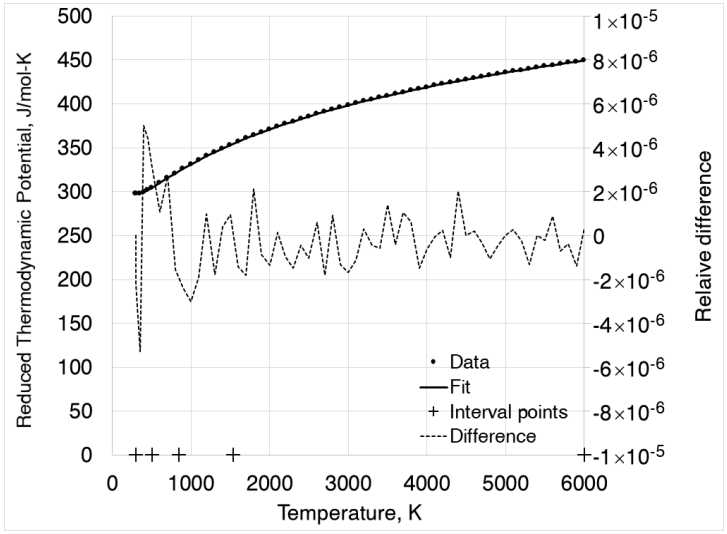
Potassium Borate (KBO_2) – B1K1O2(g)

MW = 81.91 g/mol – $\Delta_f H^o$(298.15 K) = -674.042 kJ/mol – S^o(298.15 K) = 297.408 J/mol-K – p^o = 0.1 MPa



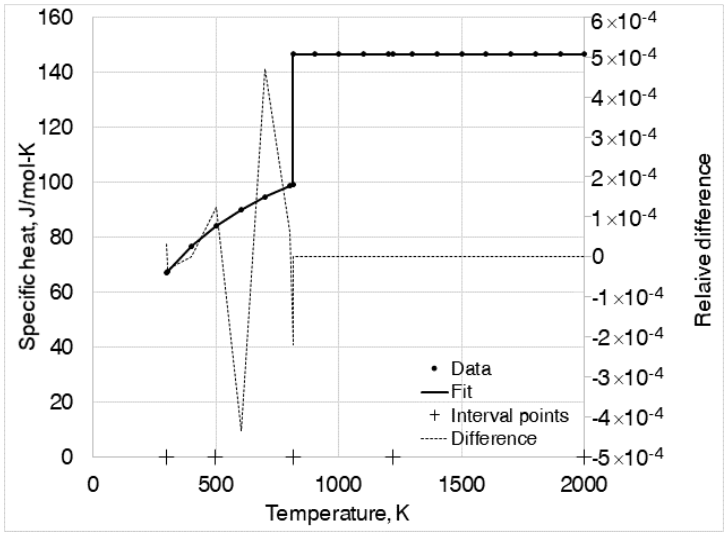


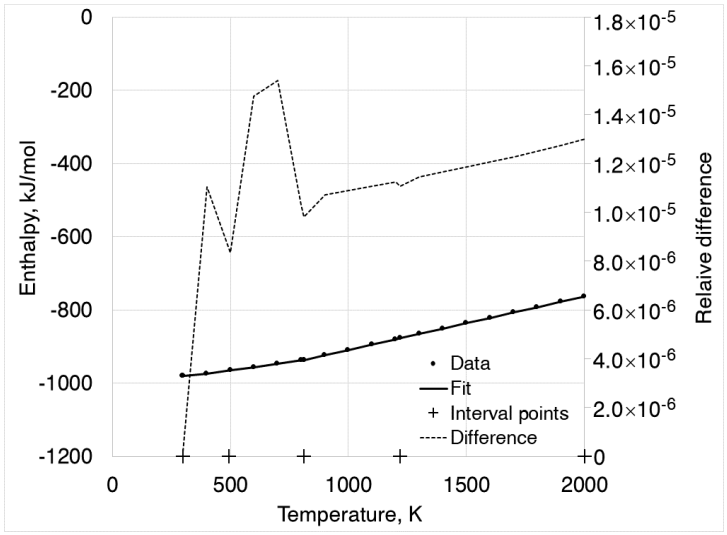


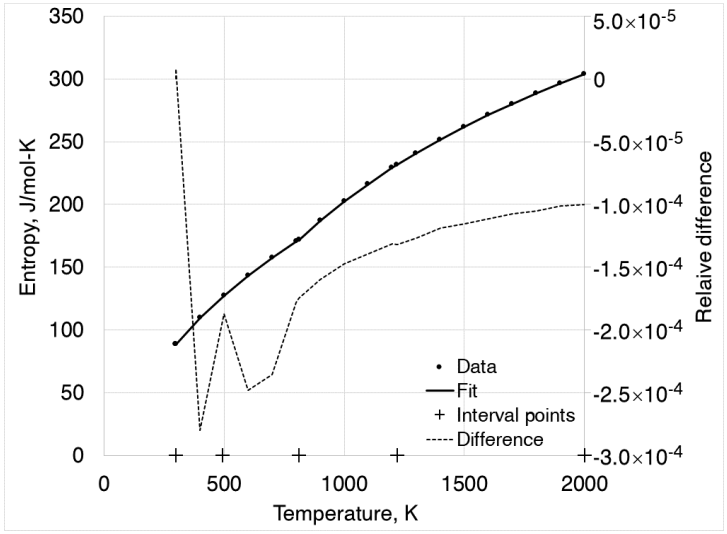


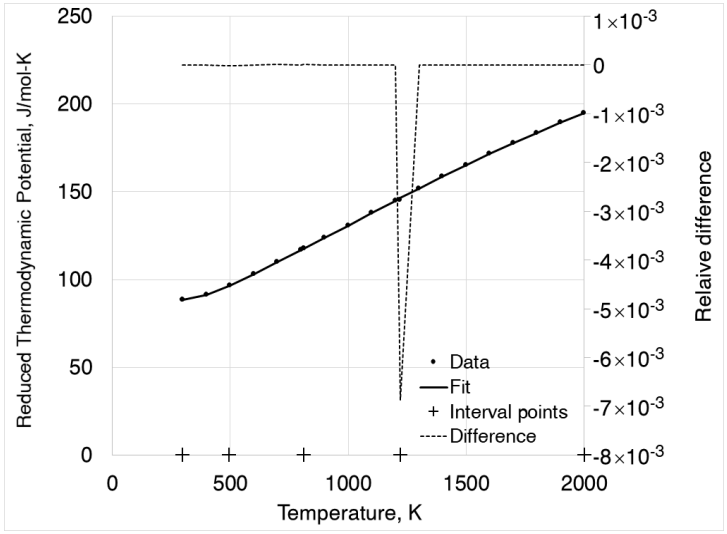
Potassium Borate (KBO_2) – B1K1O2(l)

MW = 81.91 g/mol – $\Delta_f H^o$(298.15 K) = -980.73 kJ/mol – S^o(298.15 K) = 88.455 J/mol-K – p^o = 0.1 MPa



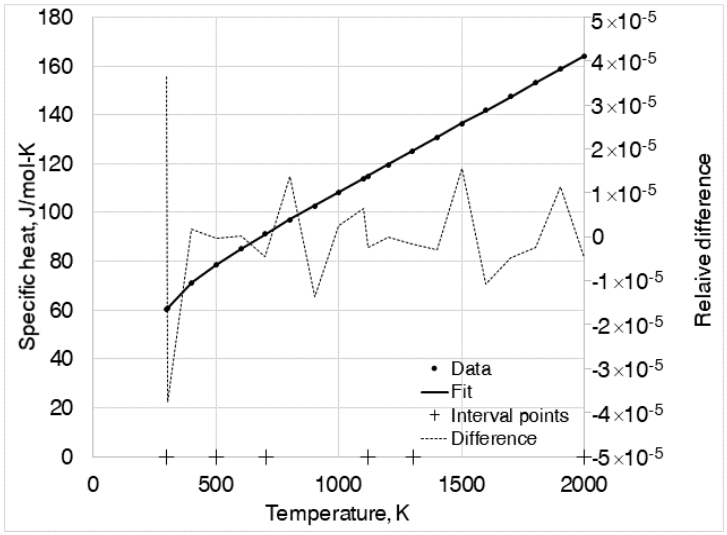


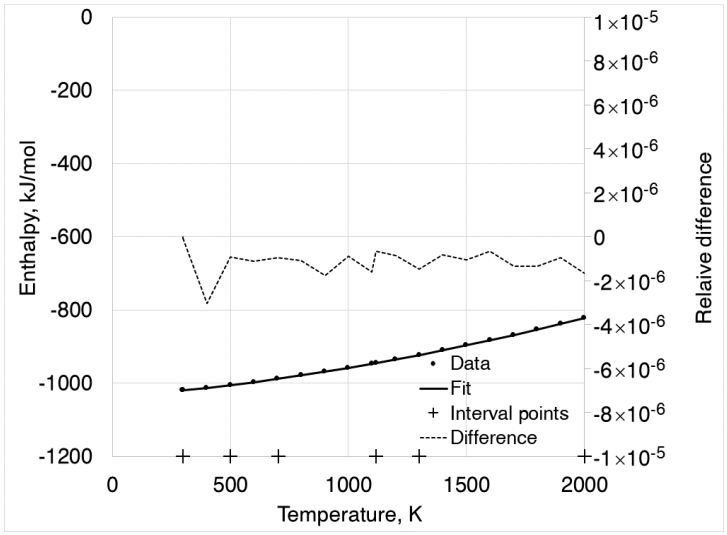


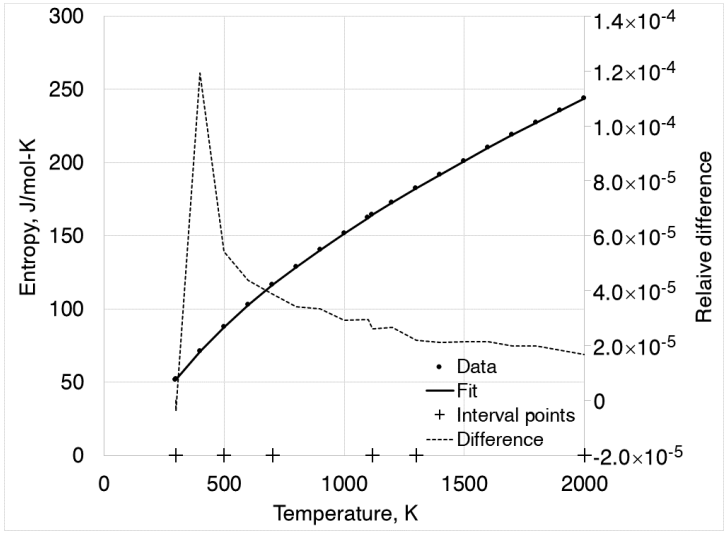


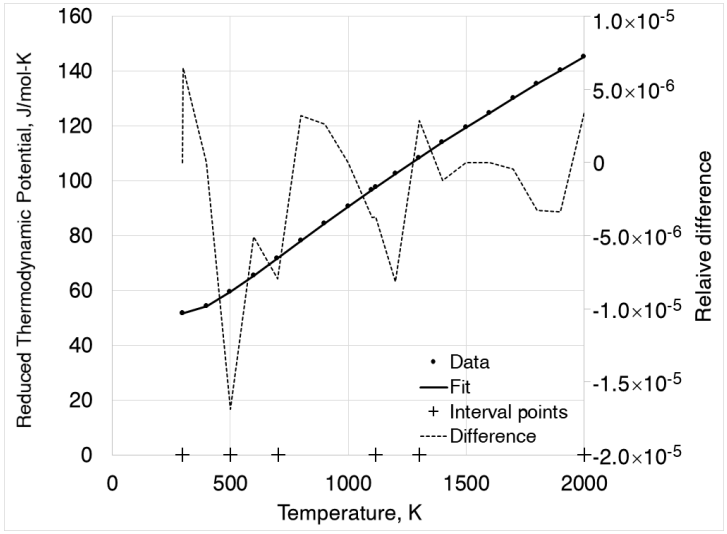
Lithium Borate ($LiBO_2$) – B1Li1O2(cr)

MW = 49.75 g/mol – $\Delta_f H^o$(298.15 K) = -1019.222 kJ/mol – S^o(298.15 K) = 51.731 J/mol-K – p^o = 0.1 MPa



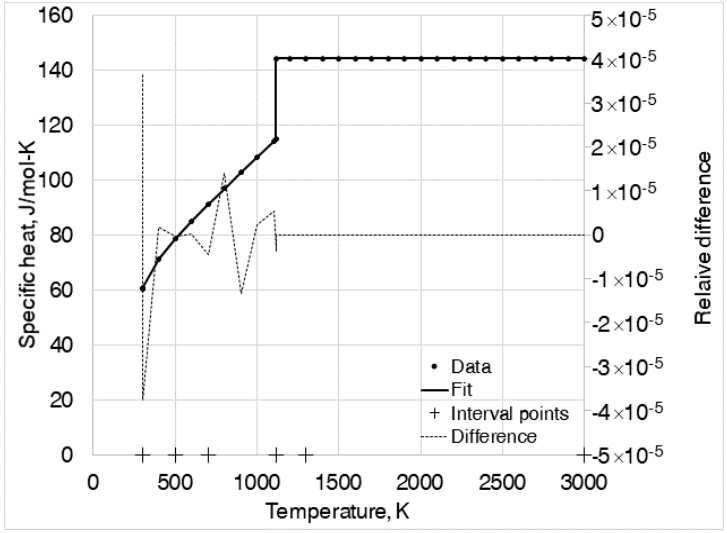


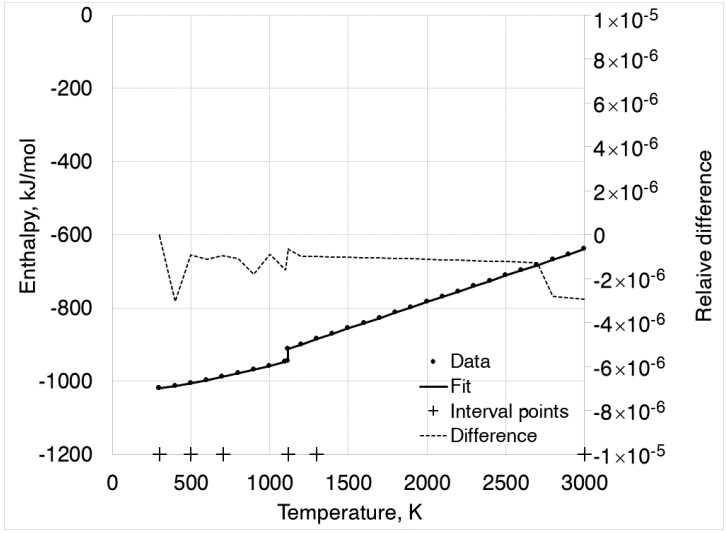


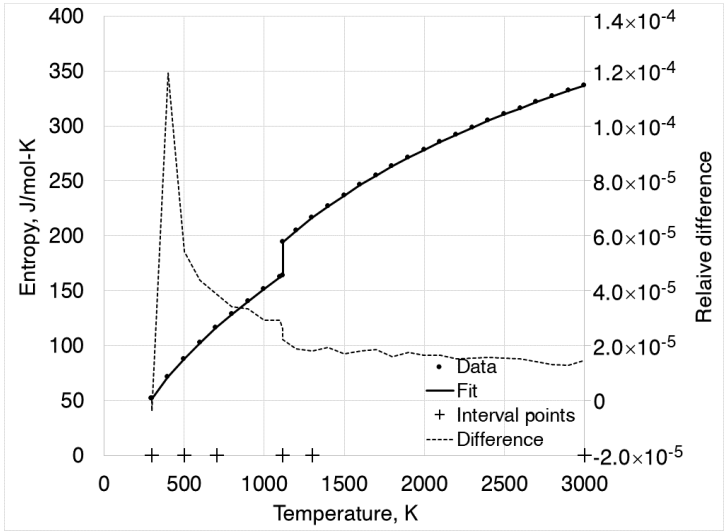


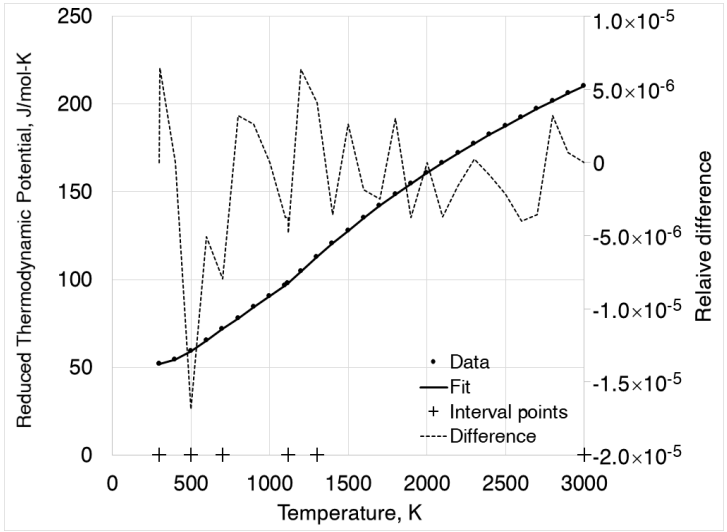
Lithium Borate ($LiBO_2$) – B1Li1O2(cr,l)

MW = 49.75 g/mol – $\Delta_f H^o$(298.15 K) = -1019.222 kJ/mol – S^o(298.15 K) = 51.731 J/mol-K – p^o = 0.1 MPa



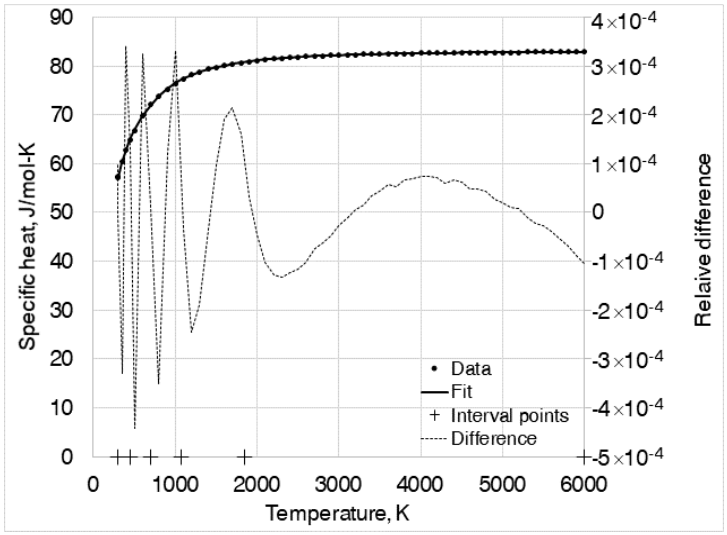


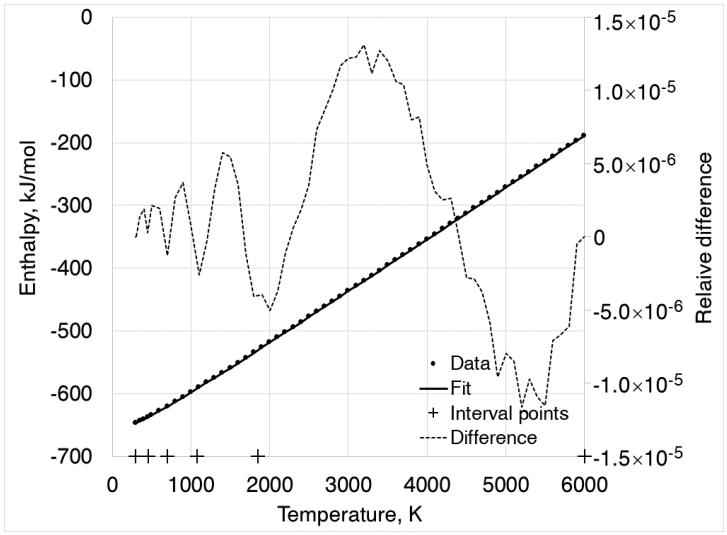


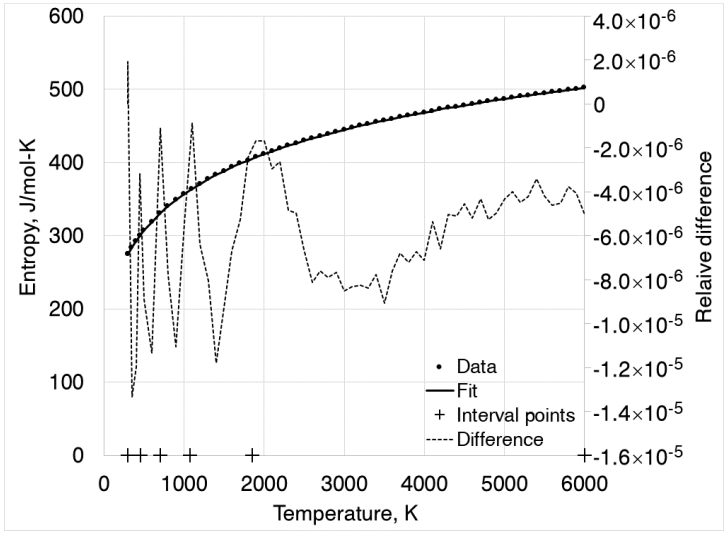


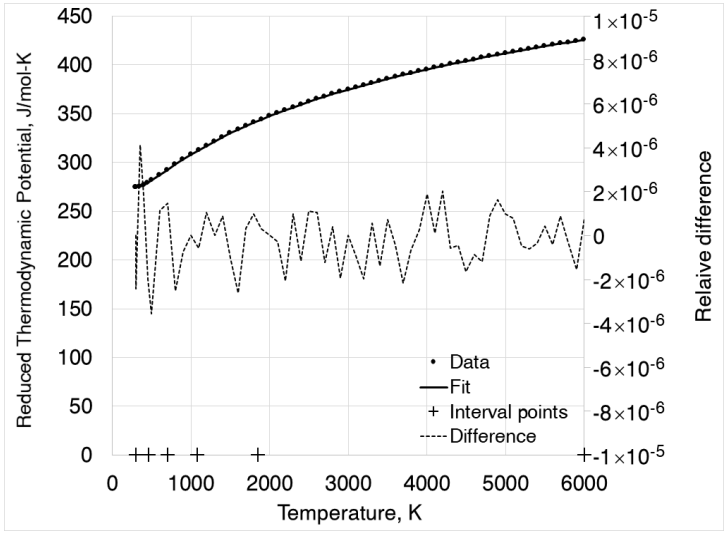
Lithium Borate ($LiBO_2$) – B1Li1O2(g)

MW = 49.75 g/mol – $\Delta_f H^o$(298.15 K) = -646.846 kJ/mol – S^o(298.15 K) = 274.727 J/mol-K – p^o = 0.1 MPa



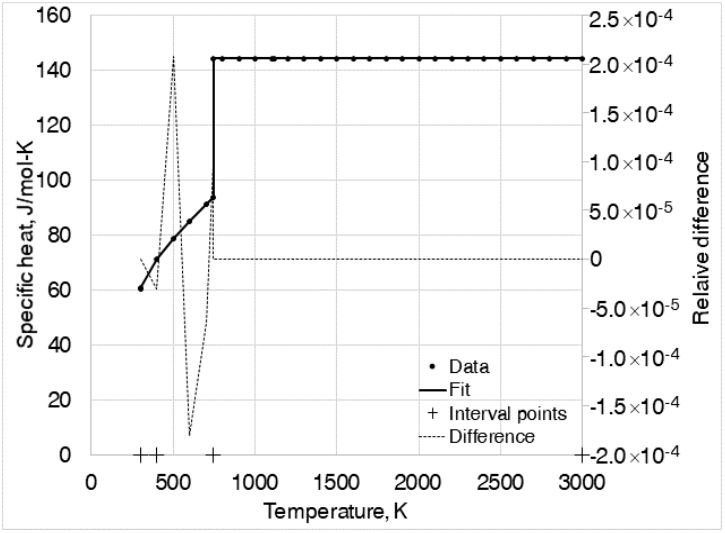


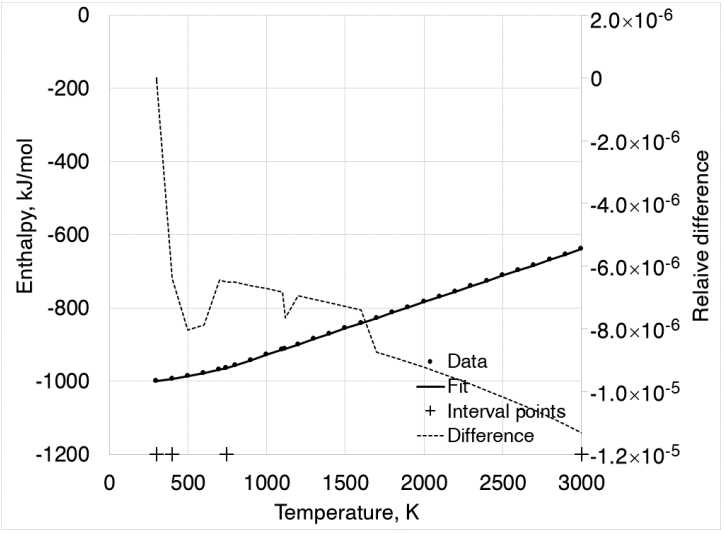


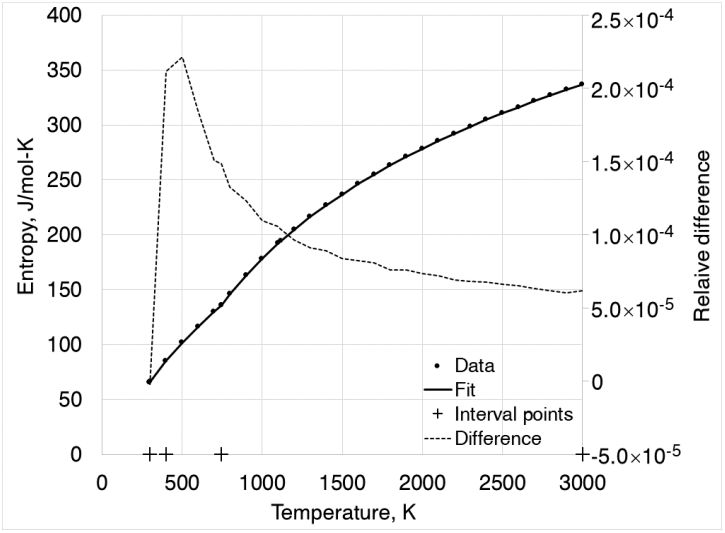


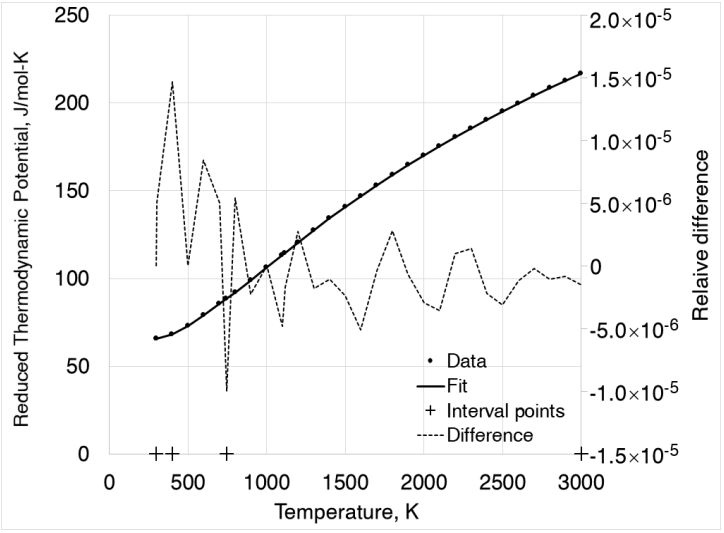
Lithium Borate ($LiBO_2$) – B1Li1O2(l)

MW = 49.75 g/mol – $\Delta_f H^o$(298.15 K) = -1000.259 kJ/mol – S^o(298.15 K) = 65.547 J/mol-K – p^o = 0.1 MPa



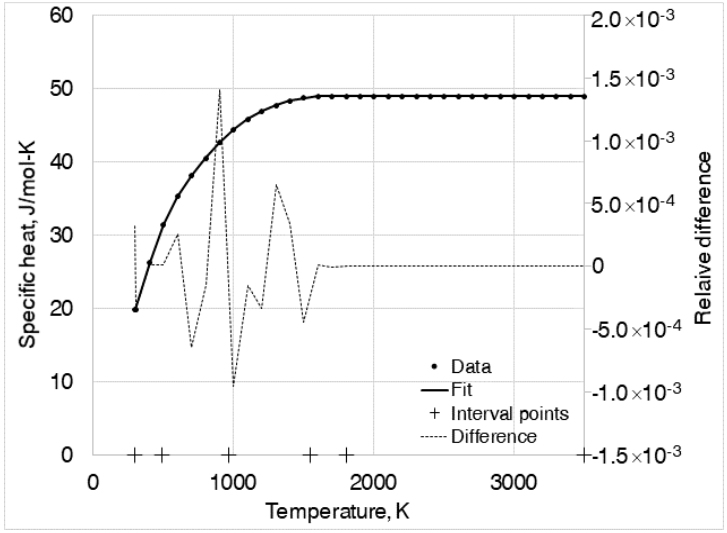


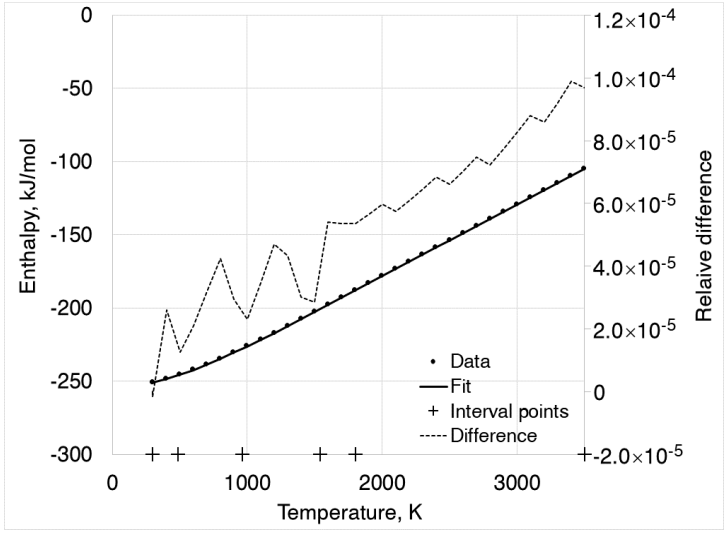


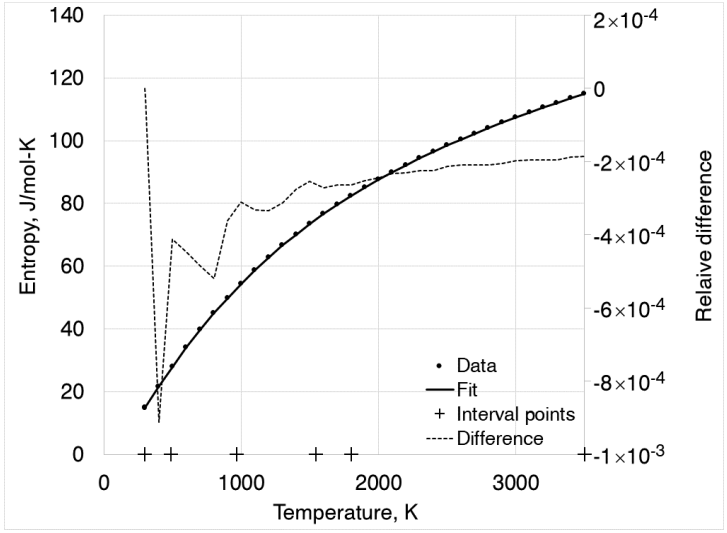


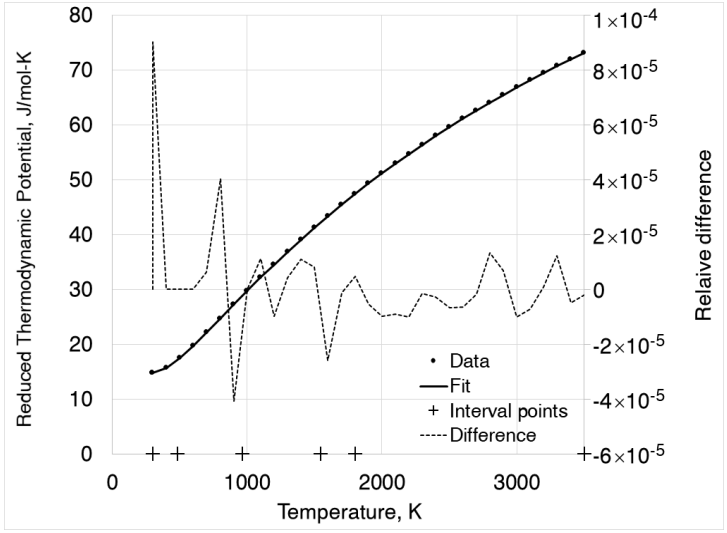
Boron Nitride (BN) – B1N1(cr)

MW = 24.82 g/mol – $\Delta_f H^o$(298.15 K) = -250.914 kJ/mol – S^o(298.15 K) = 14.795 J/mol-K – p^o = 0.1 MPa



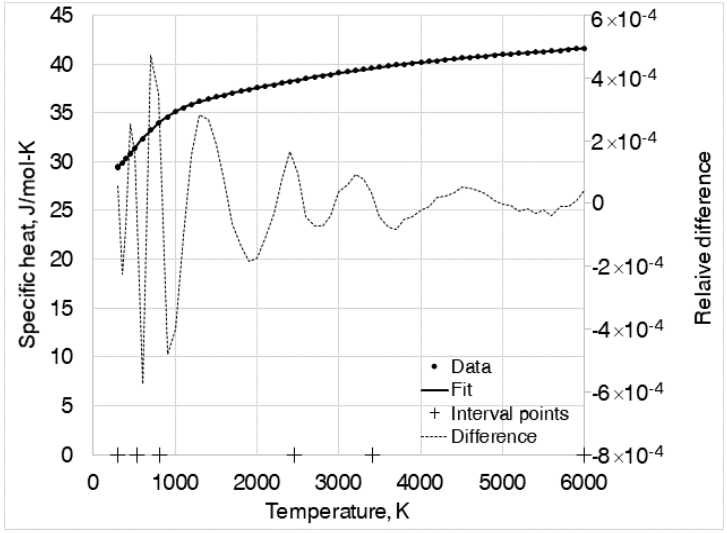


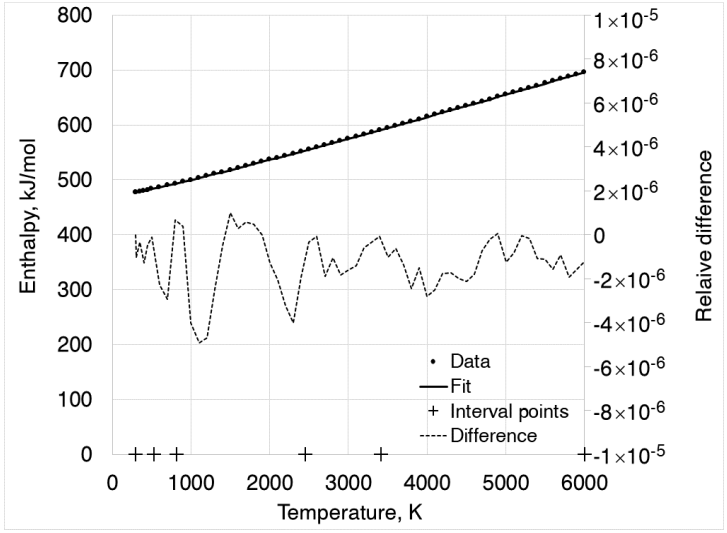


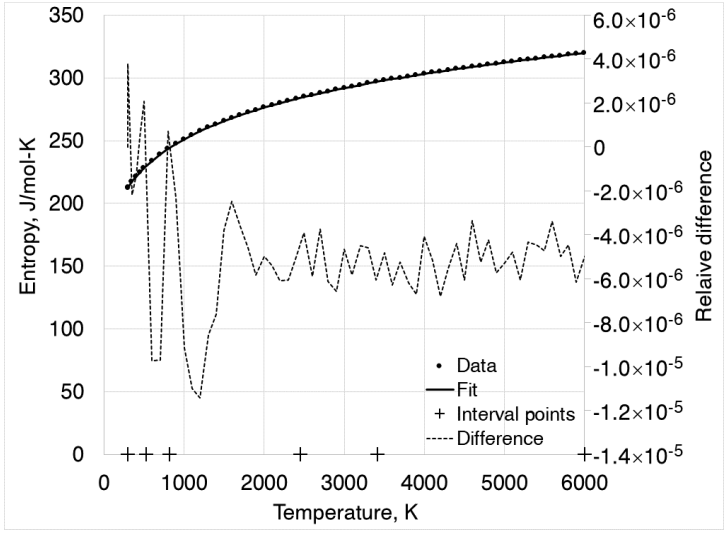


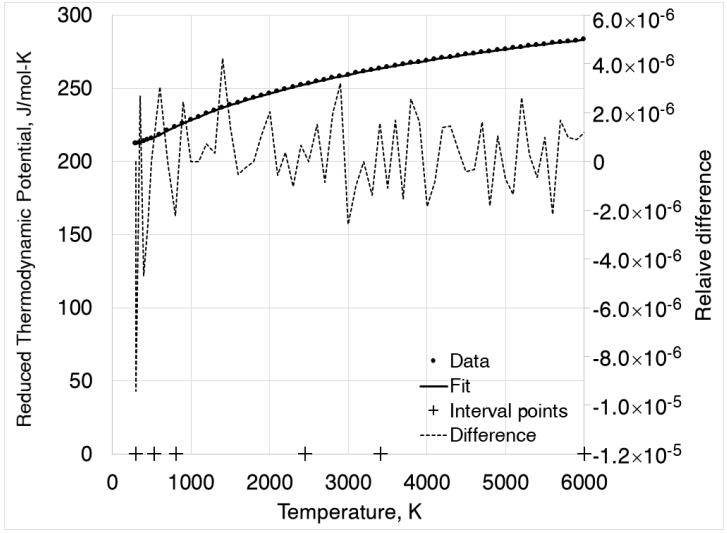
Boron Nitride (BN) – B1N1(g)

MW = 24.82 g/mol – $\Delta_f H^o$(298.15 K) = 476.976 kJ/mol – S^o(298.15 K) = 212.358 J/mol-K – p^o = 0.1 MPa



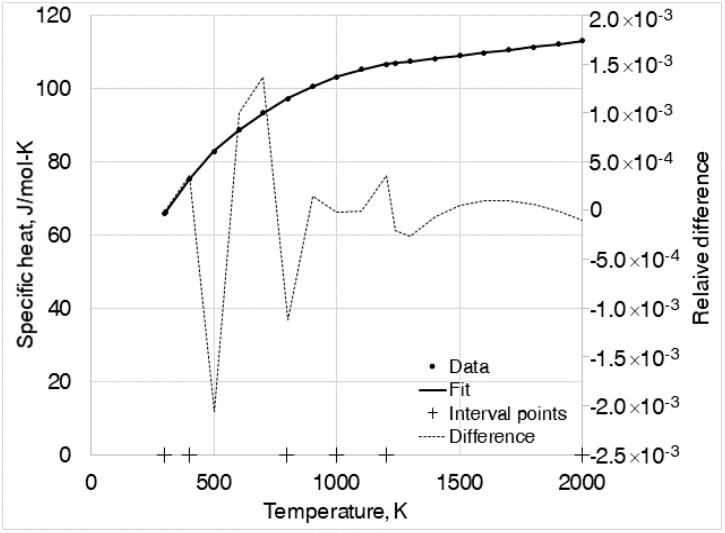


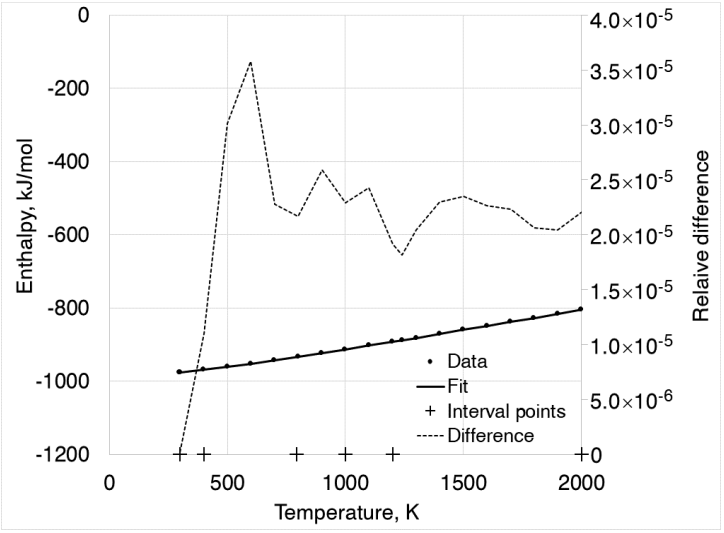


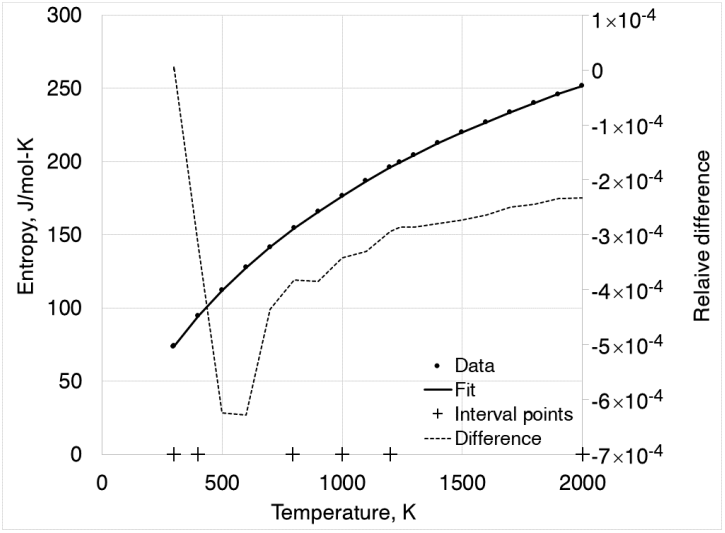


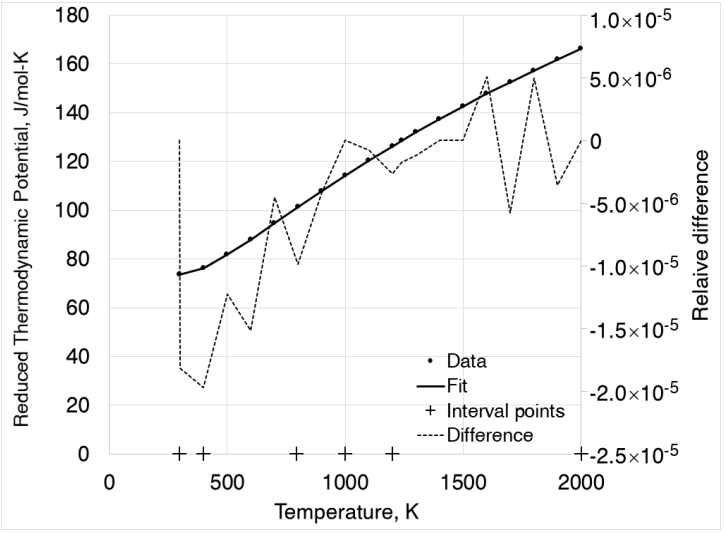
Sodium Borate ($NaBO_2$) – B1Na1O2(cr)

MW = 65.8 g/mol – $\Delta_f H^o$(298.15 K) = -975.709 kJ/mol – S^o(298.15 K) = 73.538 J/mol-K – p^o = 0.1 MPa



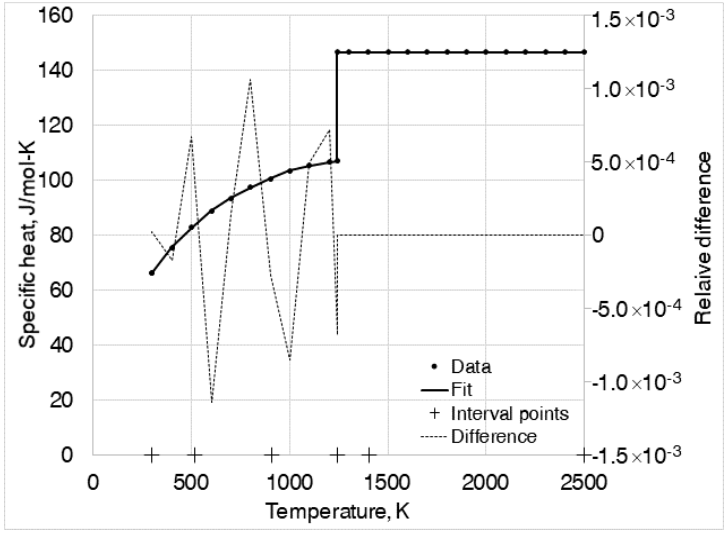


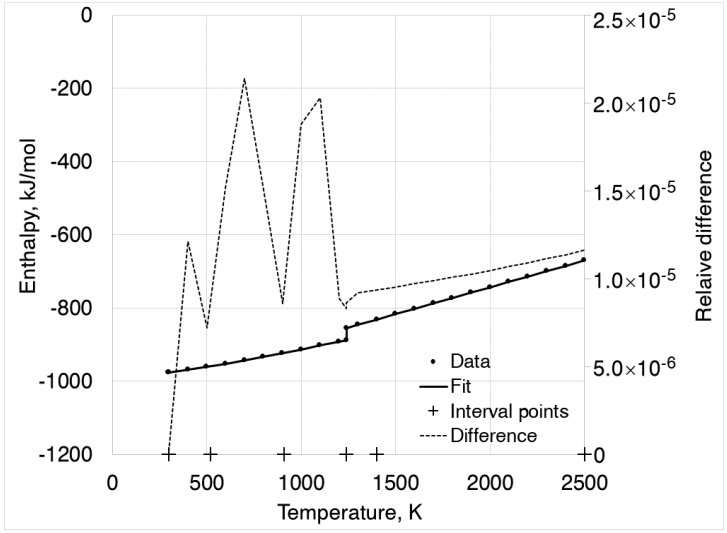


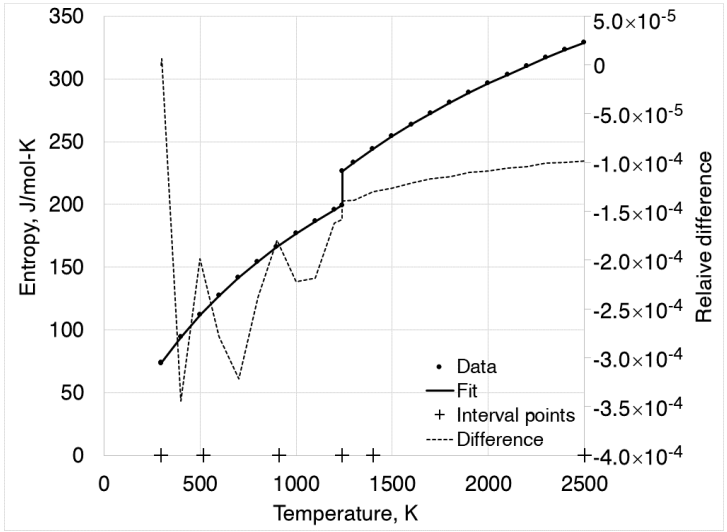


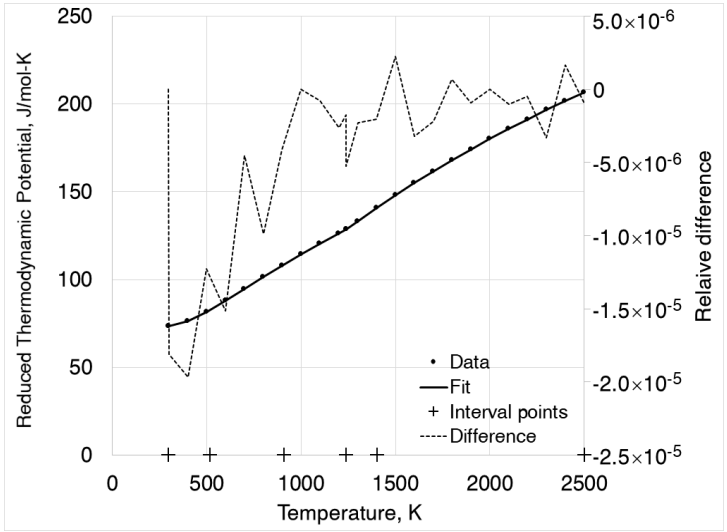
Sodium Borate ($NaBO_2$) – B1Na1O2(cr,l)

MW = 65.8 g/mol – $\Delta_f H^o$(298.15 K) = -975.709 kJ/mol – S^o(298.15 K) = 73.538 J/mol-K – p^o = 0.1 MPa



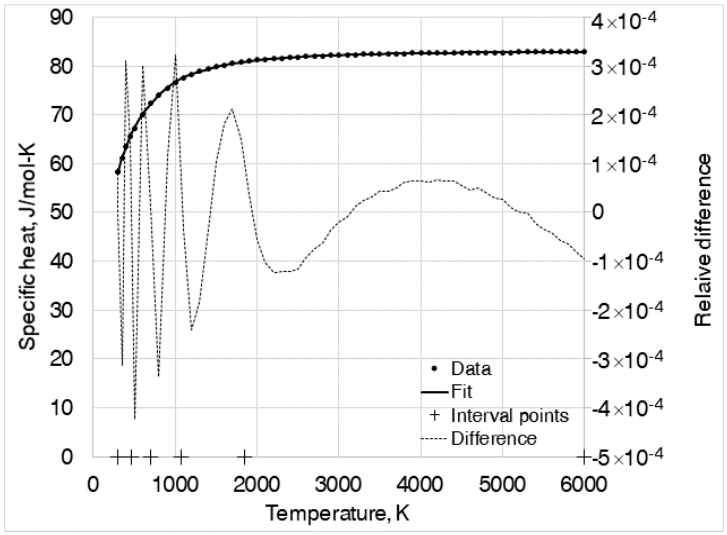


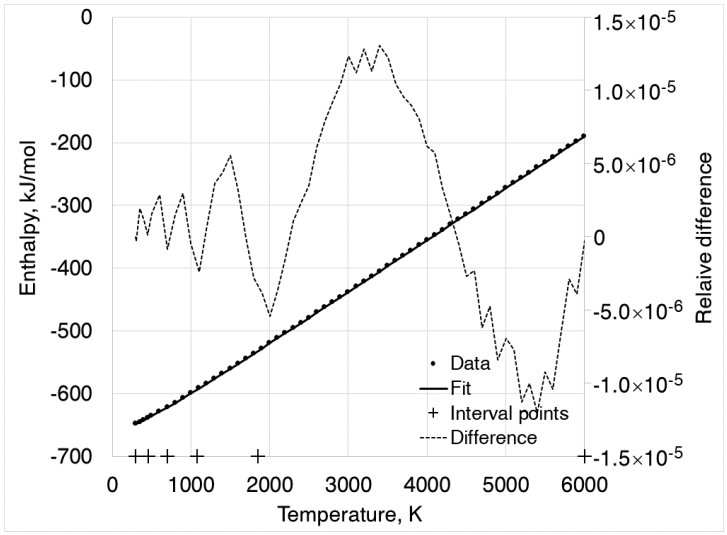


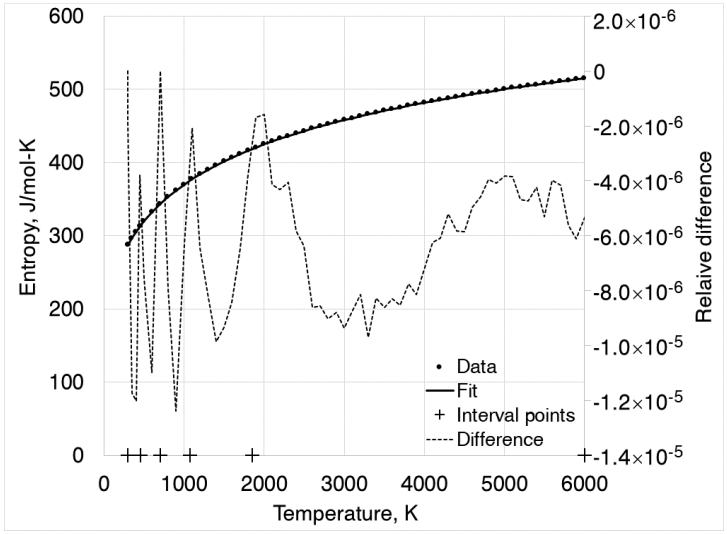


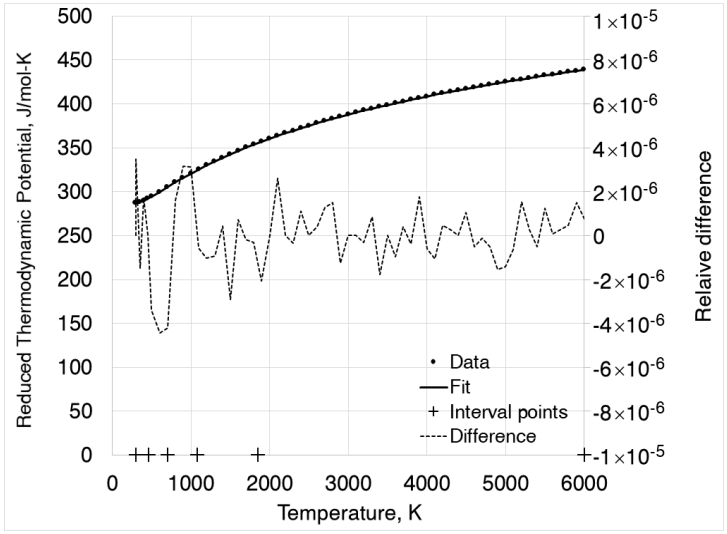
Sodium Borate ($NaBO_2$) – B1Na1O2(g)

MW = 65.8 g/mol – $\Delta_f H^o$(298.15 K) = -648.52 kJ/mol – S^o(298.15 K) = 287.248 J/mol-K – p^o = 0.1 MPa



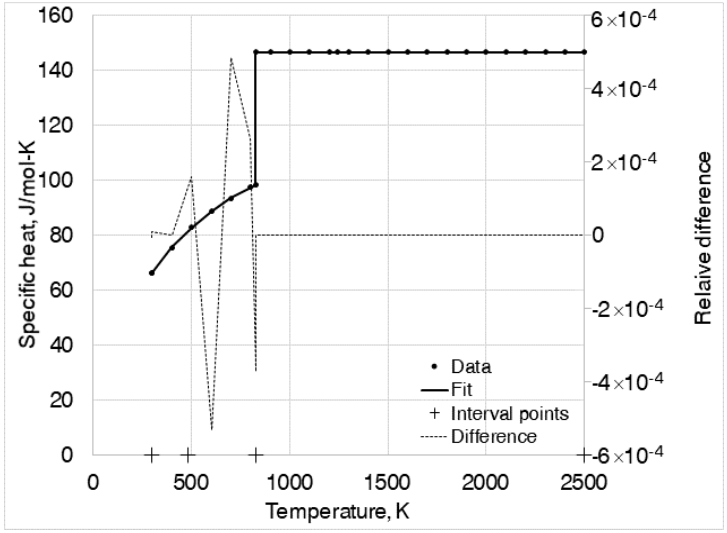


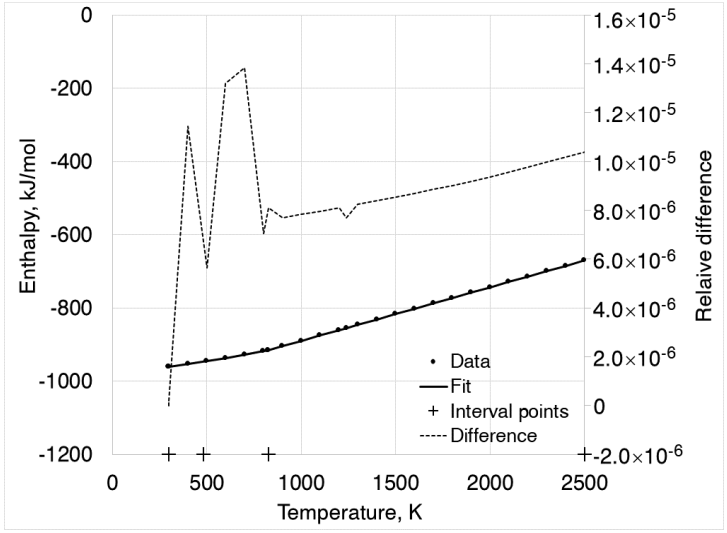


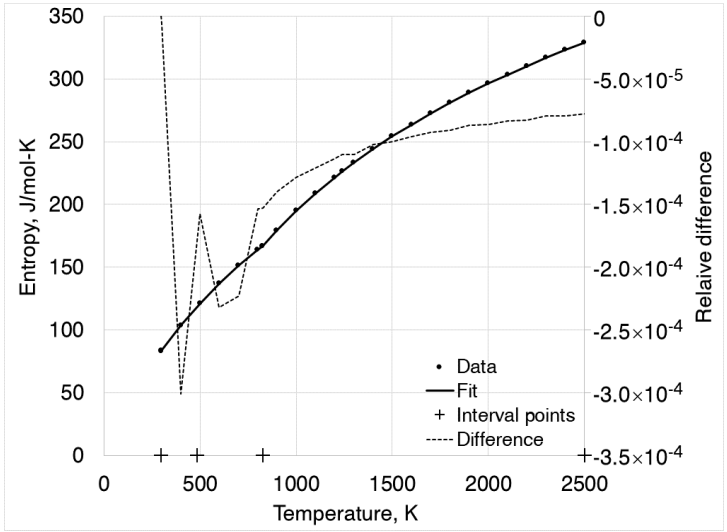


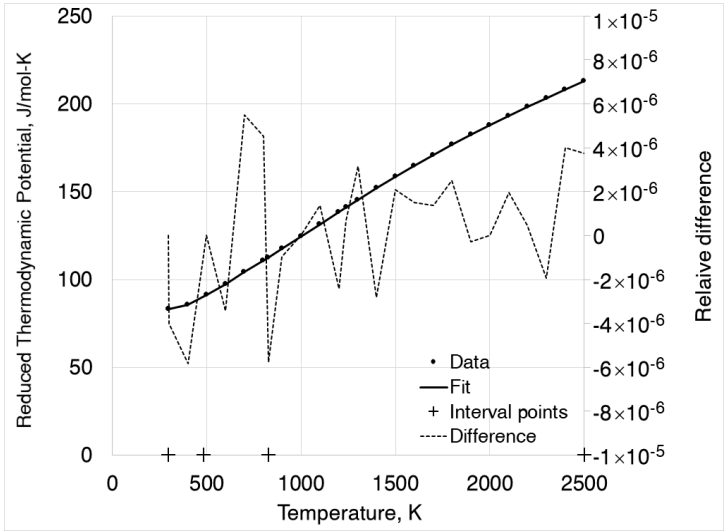
Sodium Borate ($NaBO_2$) – B1Na1O2(l)

MW = 65.8 g/mol – $\Delta_f H^o$(298.15 K) = -960.019 kJ/mol – S^o(298.15 K) = 82.958 J/mol-K – p^o = 0.1 MPa



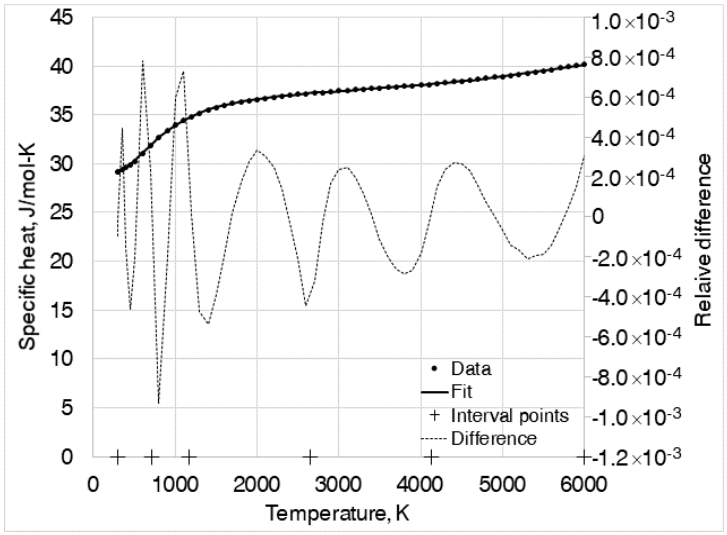


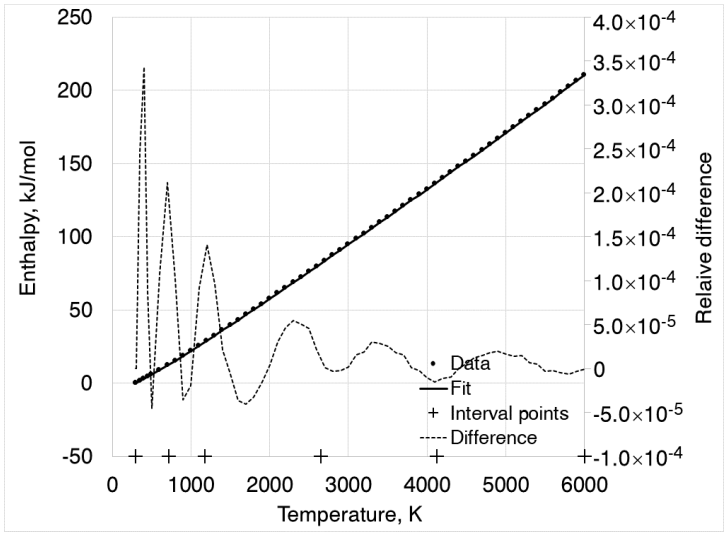


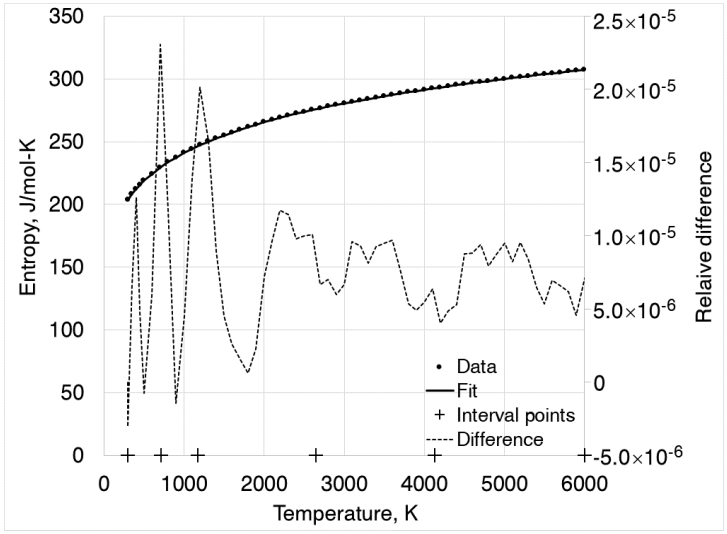


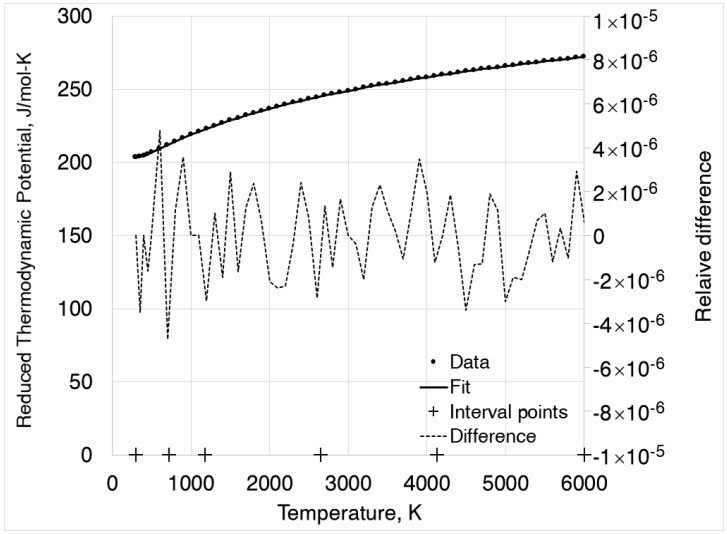
Boron Oxide (BO) – B1O1(g)

MW = 26.81 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 203.472 J/mol-K – p^o = 0.1 MPa



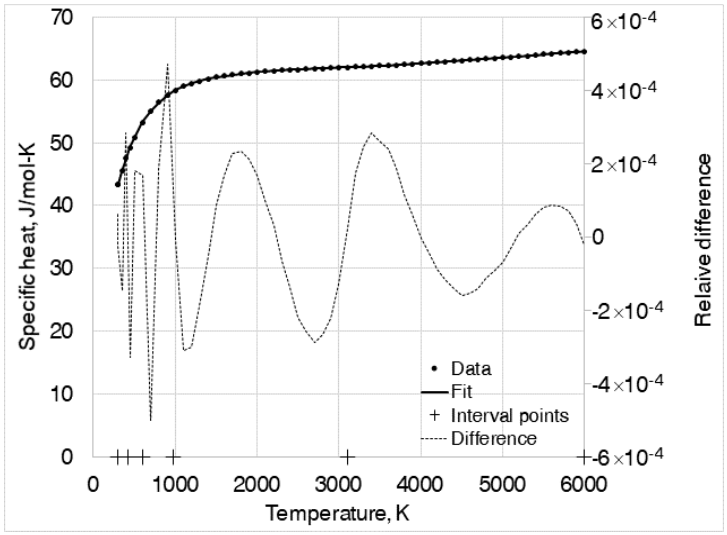


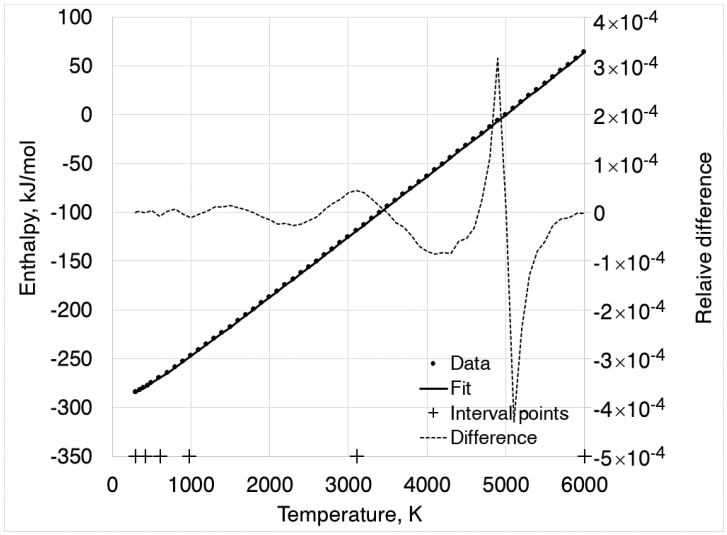


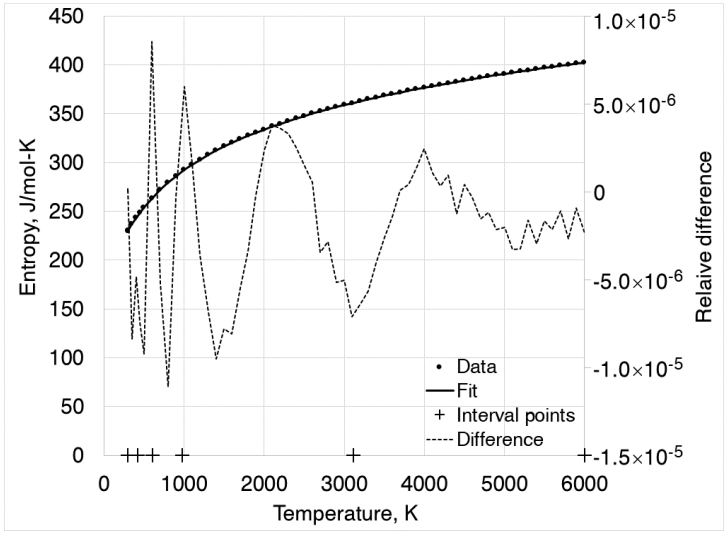


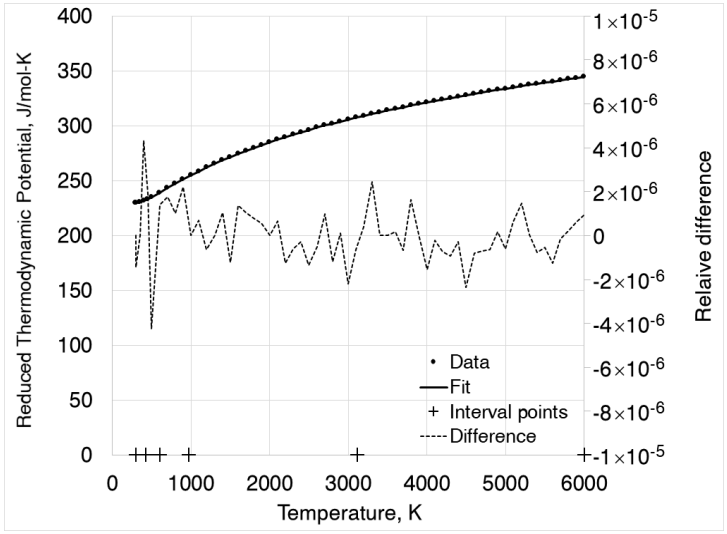
Boron Oxide (BO_2) – B1O2(g)

MW = 42.81 g/mol – $\Delta_f H^o$(298.15 K) = -284.512 kJ/mol – S^o(298.15 K) = 229.814 J/mol-K – p^o = 0.1 MPa



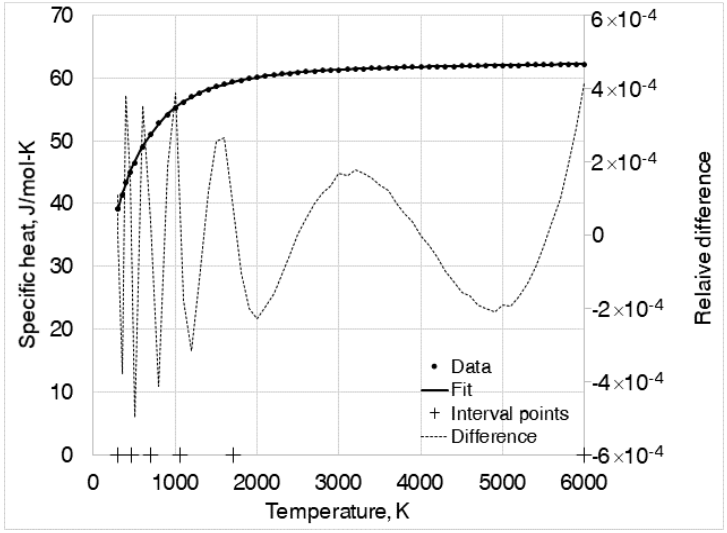


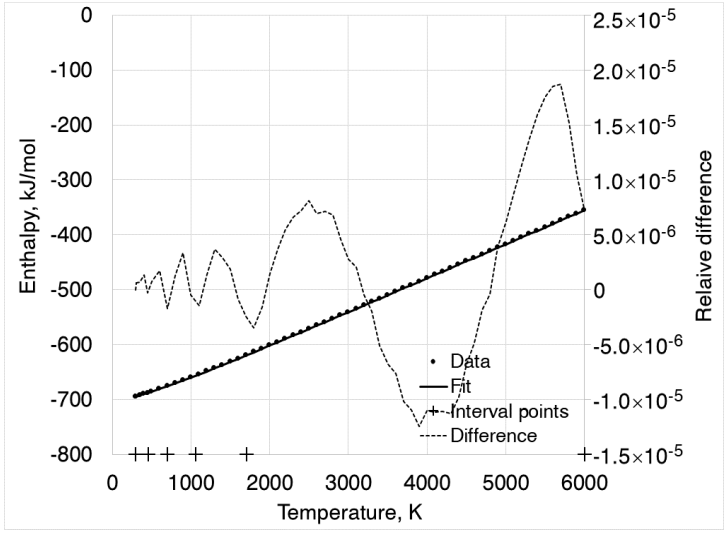


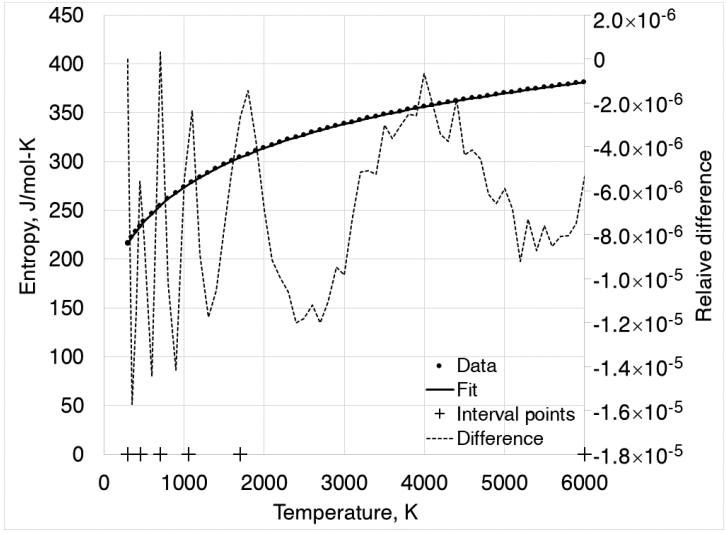


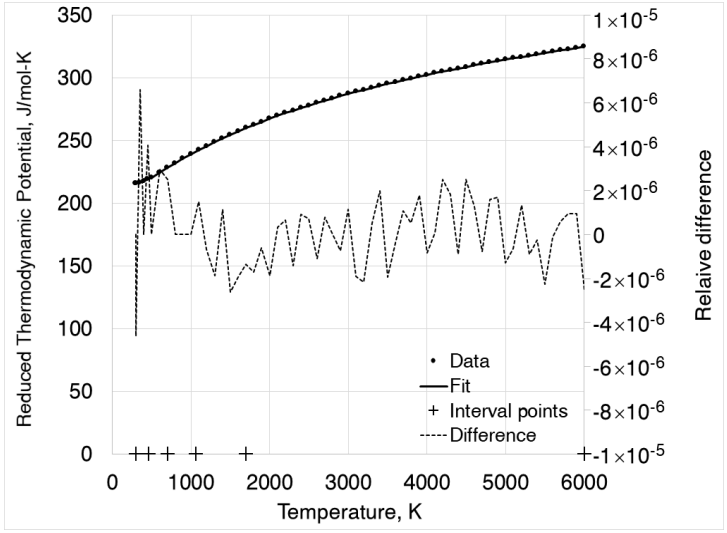
Boron Oxide, Ion (BO_2^-) – B1O2-(g)

MW = 42.81 g/mol – $\Delta_f H^o$(298.15 K) = -694.544 kJ/mol – S^o(298.15 K) = 215.922 J/mol-K – p^o = 0.1 MPa



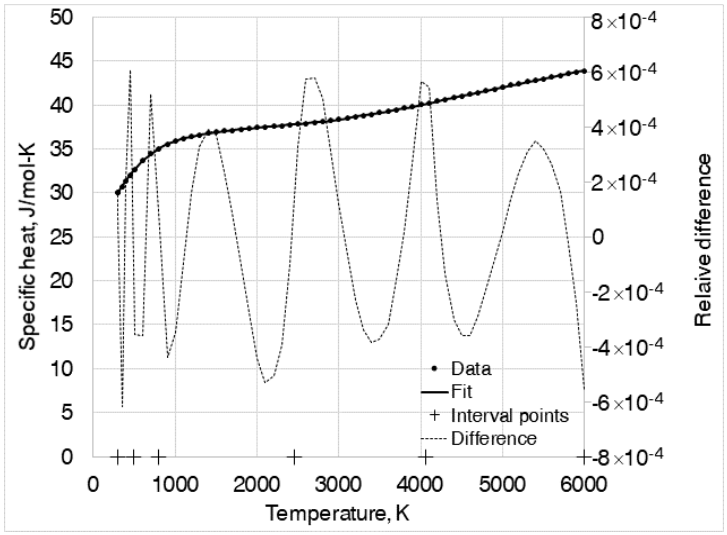


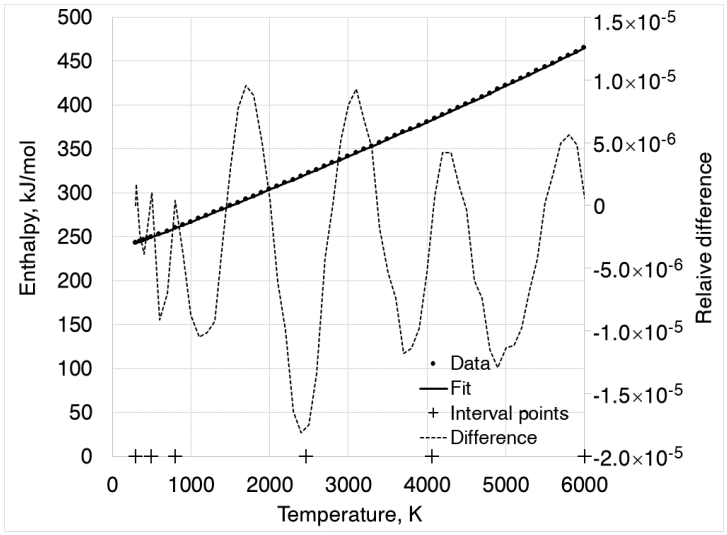


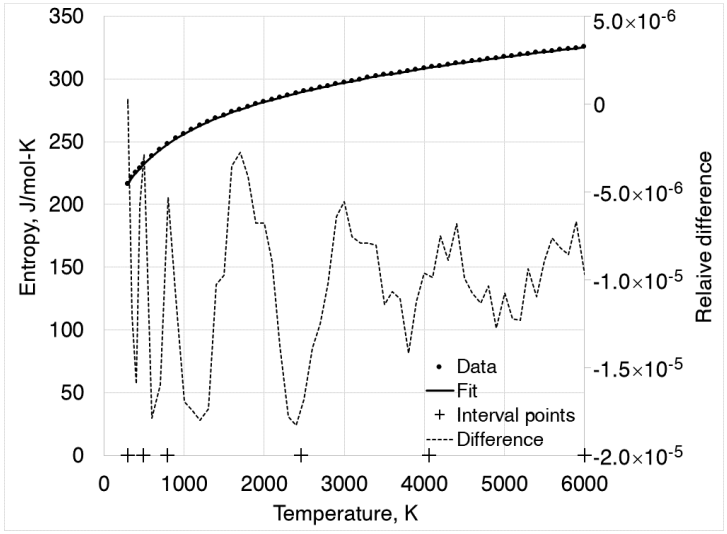


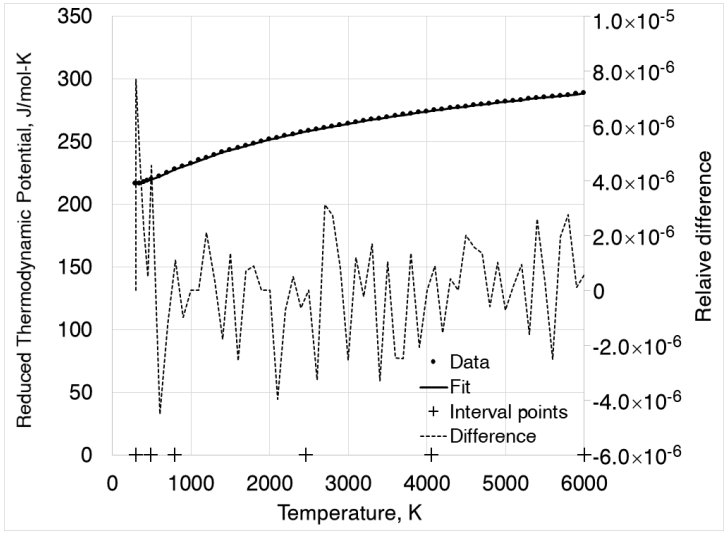
Boron Sulfide (BS) – B1S1(g)

MW = 42.87 g/mol – $\Delta_f H^o$(298.15 K) = 243.09 kJ/mol – S^o(298.15 K) = 216.19 J/mol-K – p^o = 0.1 MPa



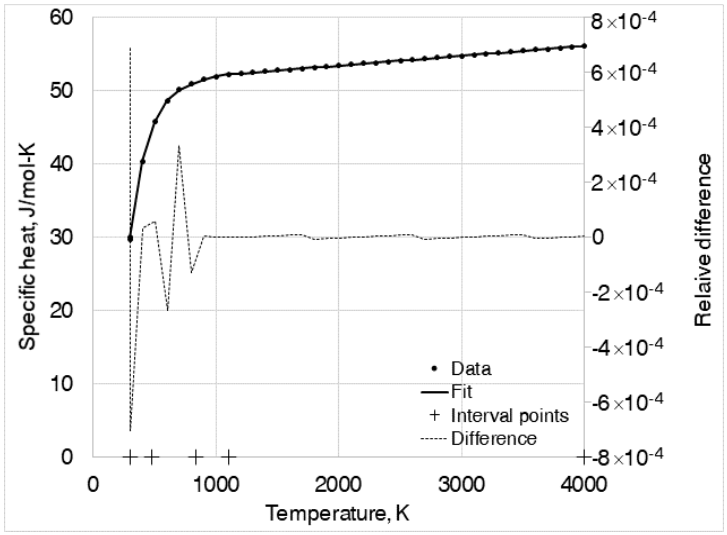


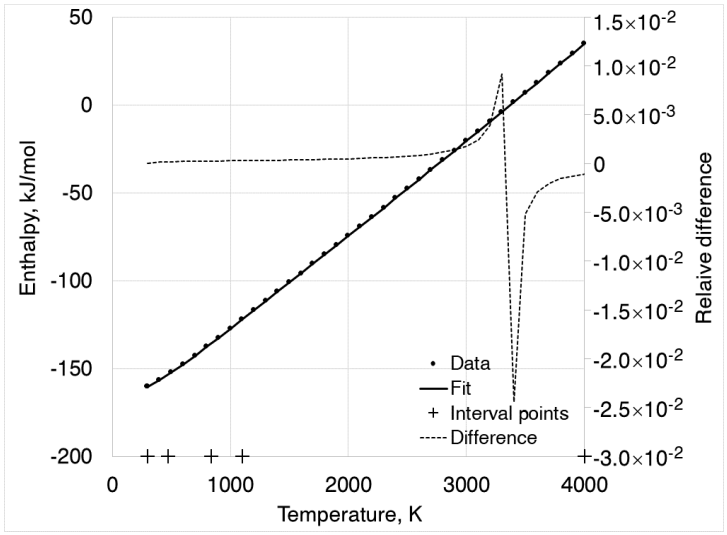


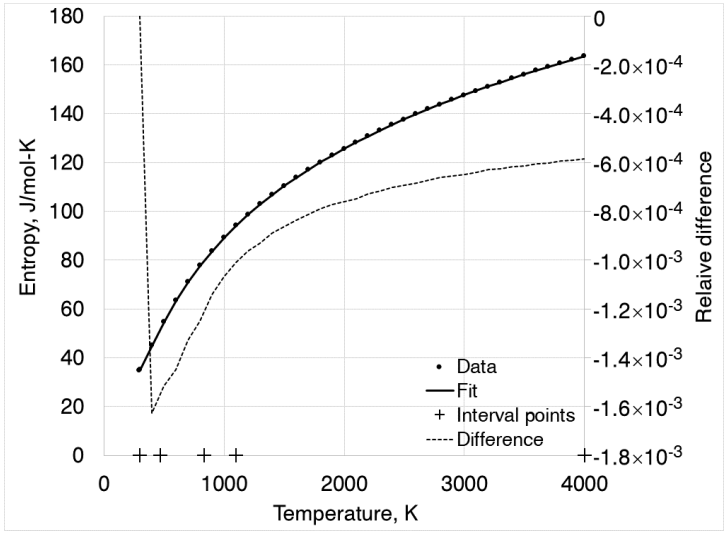


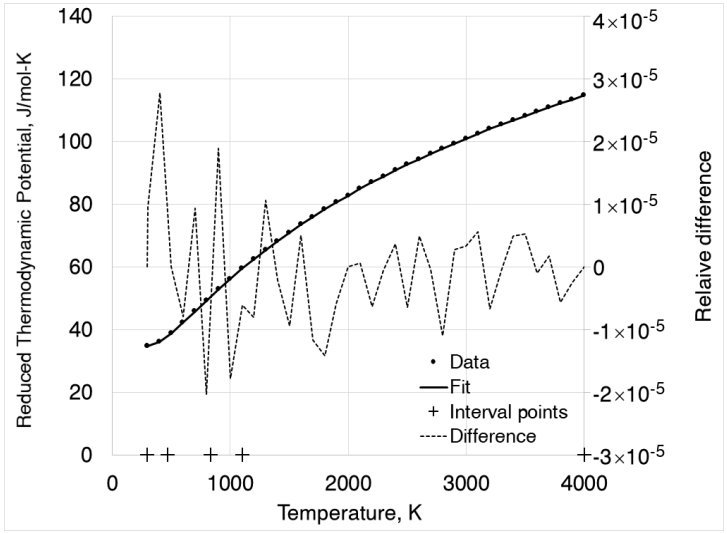
Titanium Boride (TiB) – B1Ti1(cr)

MW = 58.69 g/mol – $\Delta_f H^o$(298.15 K) = -160.247 kJ/mol – S^o(298.15 K) = 34.727 J/mol-K – p^o = 0.1 MPa



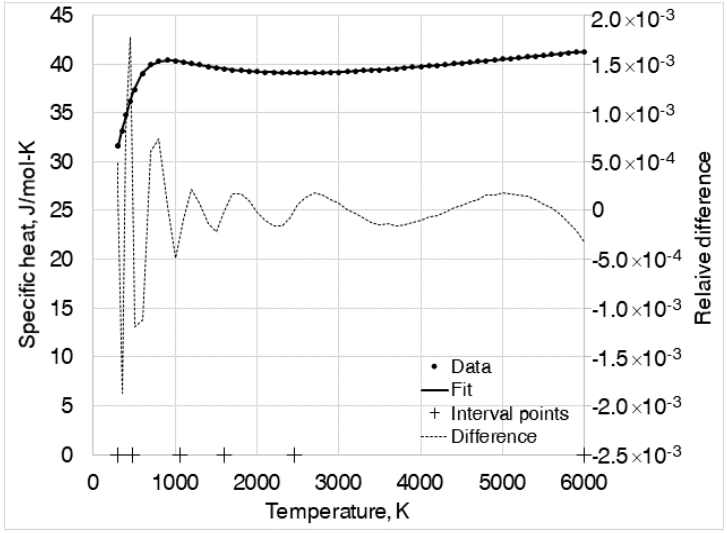


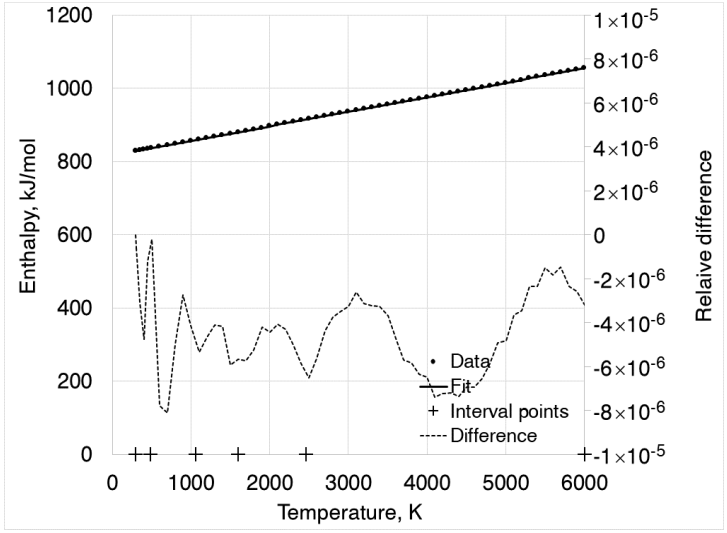


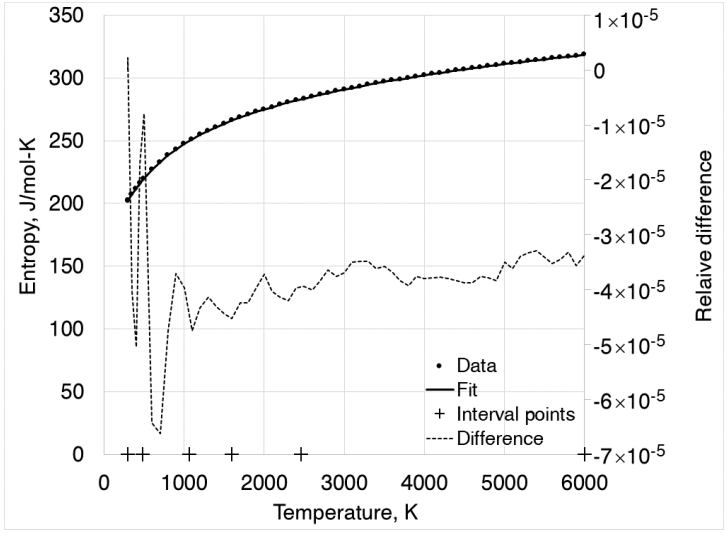


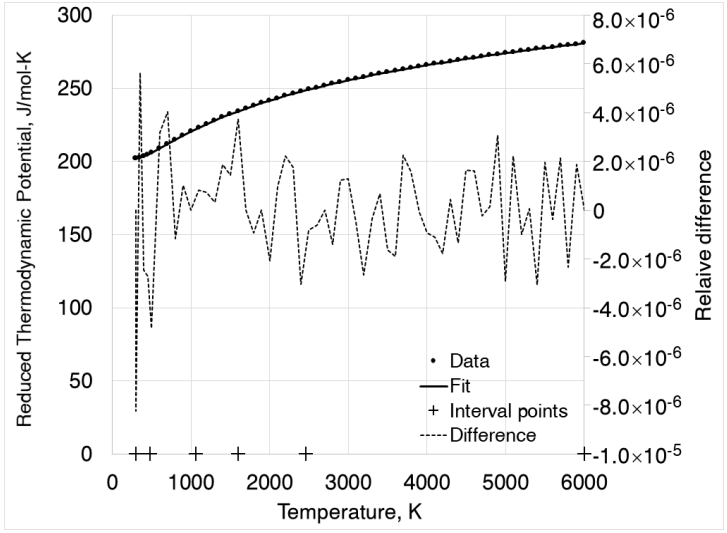
Boron (B_2) – B2(g)

MW = 21.62 g/mol – $\Delta_f H^o$(298.15 K) = 829.687 kJ/mol – S^o(298.15 K) = 202.072 J/mol-K – p^o = 0.1 MPa



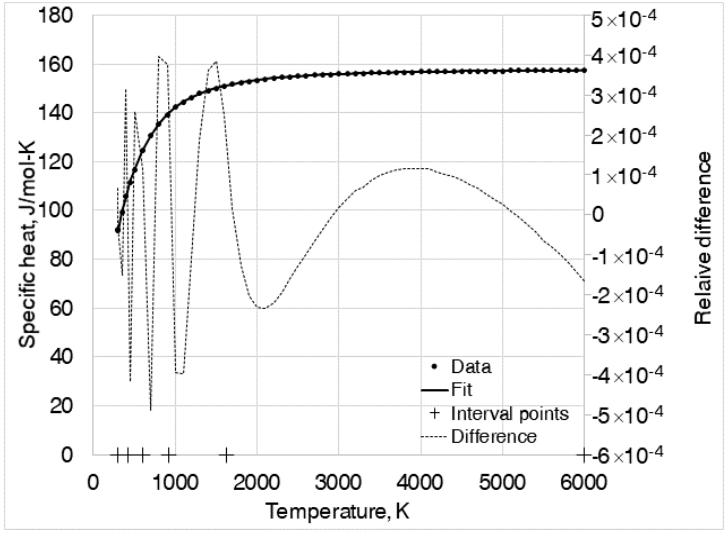


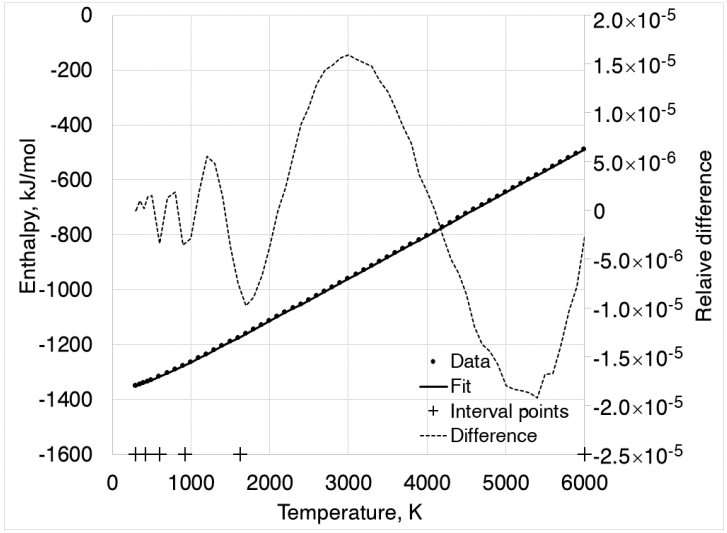


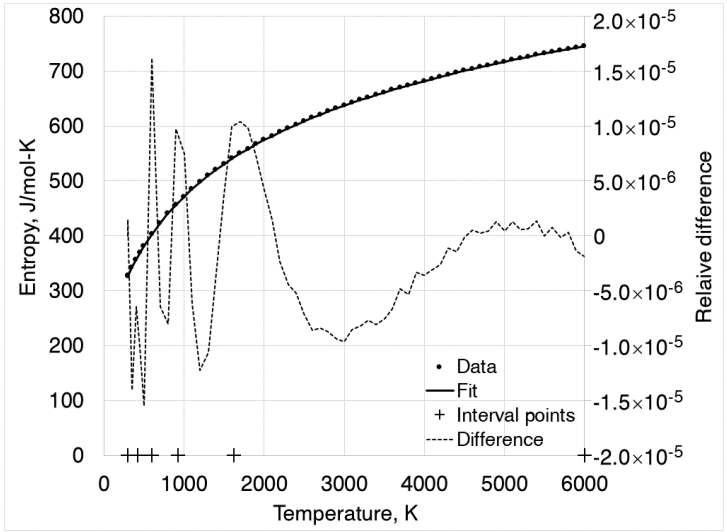


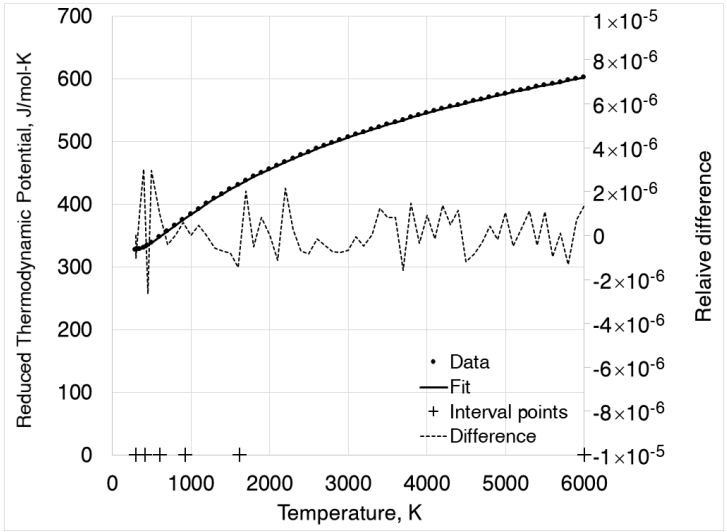
Beryllium Borate ($Be(BO_2)_2$) – B2Be1O4(g)

MW = 94.63 g/mol – $\Delta_f H^o$(298.15 K) = -1351.432 kJ/mol – S^o(298.15 K) = 326.785 J/mol-K – p^o = 0.1 MPa



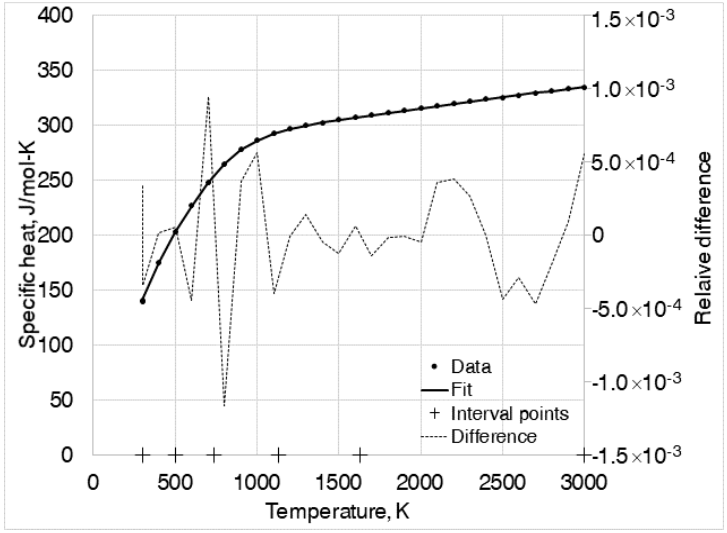


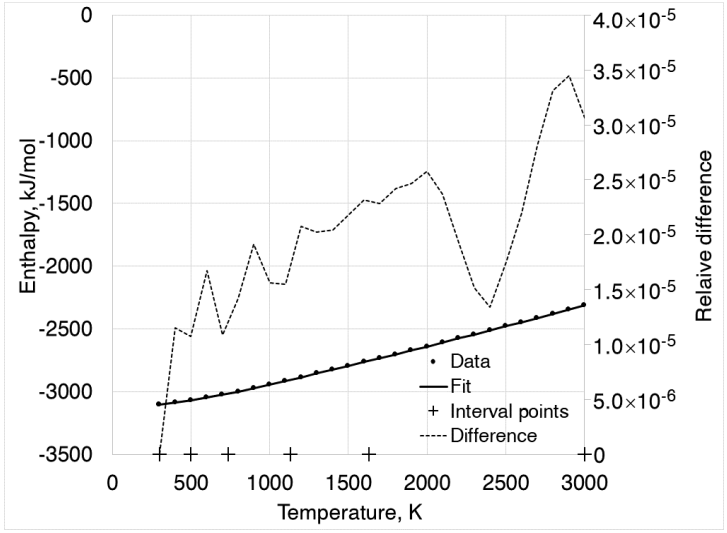


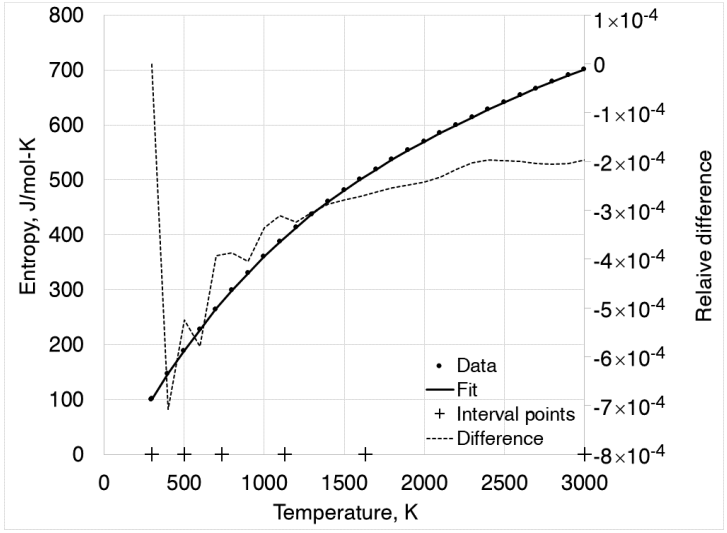


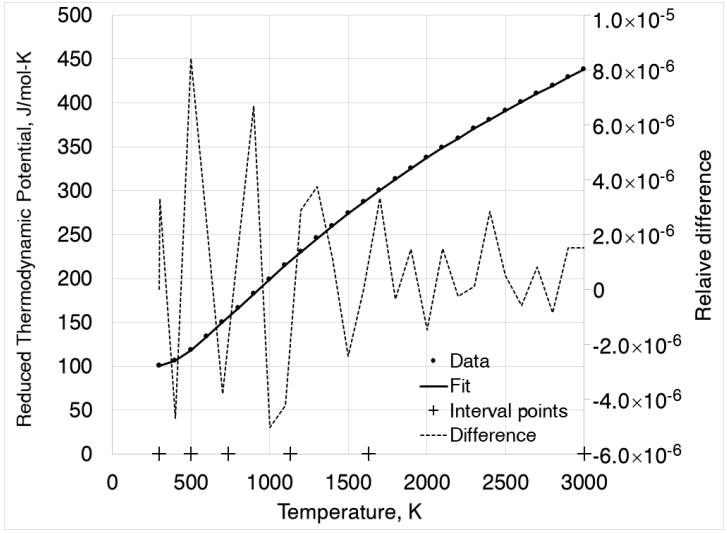
Beryllium Borate ($Be_3B_2O_6$) – B2Be3O6(cr)

MW = 144.65 g/mol – $\Delta_f H^o$(298.15 K) = -3104.361 kJ/mol – S^o(298.15 K) = 100.416 J/mol-K – p^o = 0.1 MPa



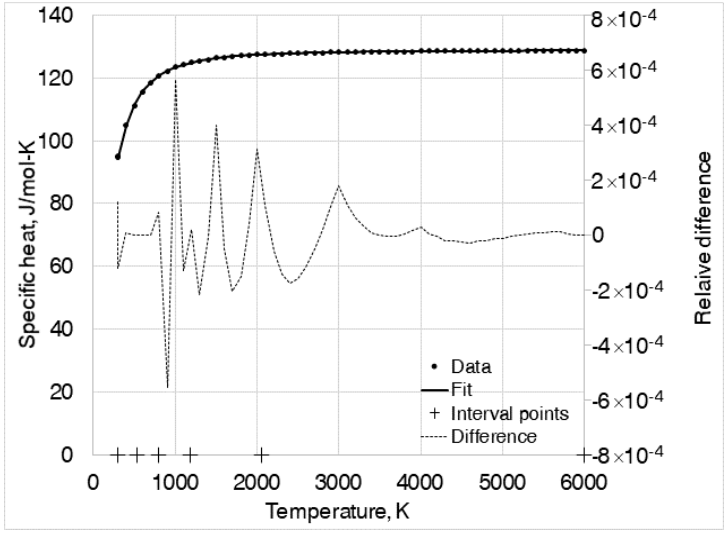


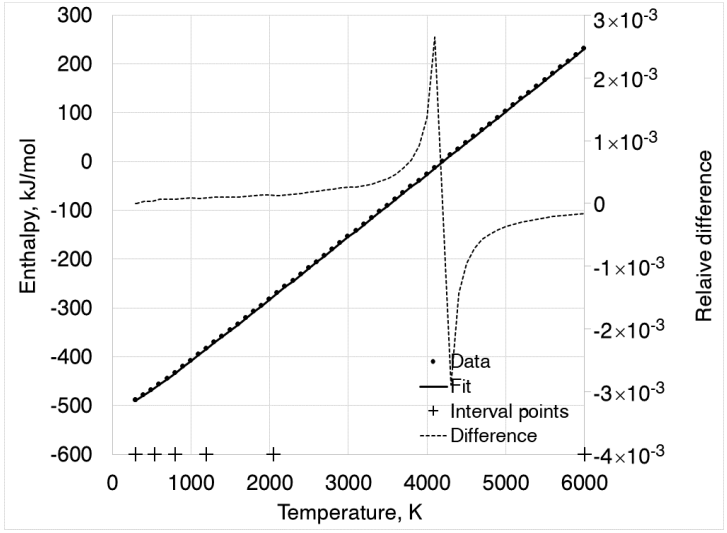


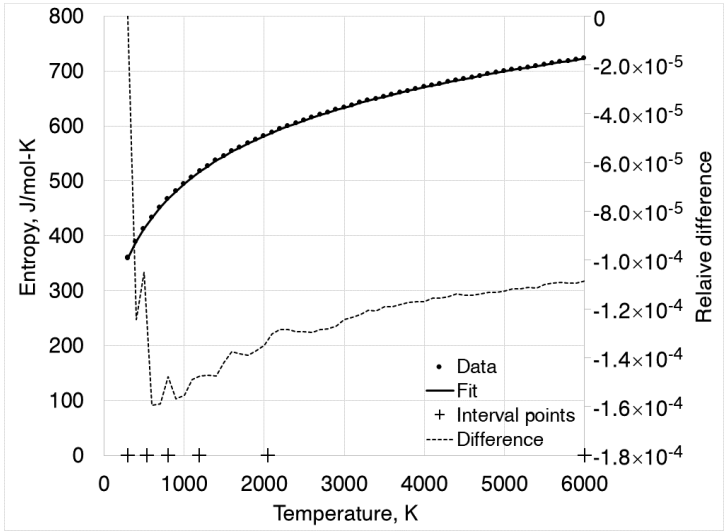


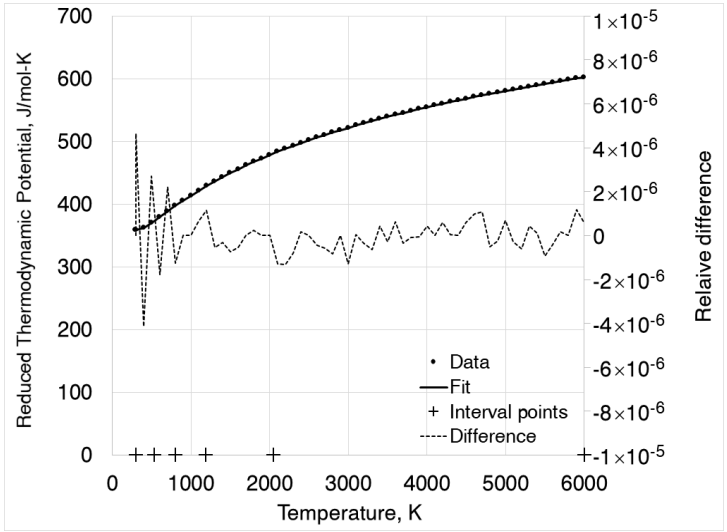
Dichloroborane ($(BCl_2)_2$) – B2Cl4(g)

MW = 163.43 g/mol – $\Delta_f H^o$(298.15 K) = -489.11 kJ/mol – S^o(298.15 K) = 358.956 J/mol-K – p^o = 0.1 MPa



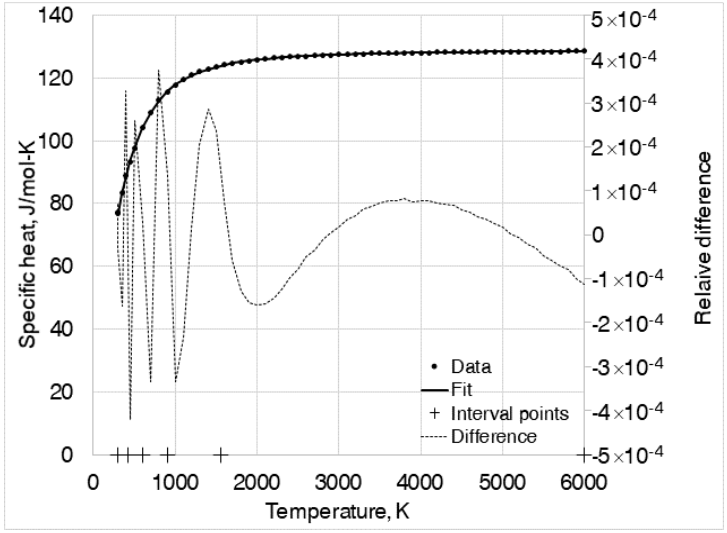


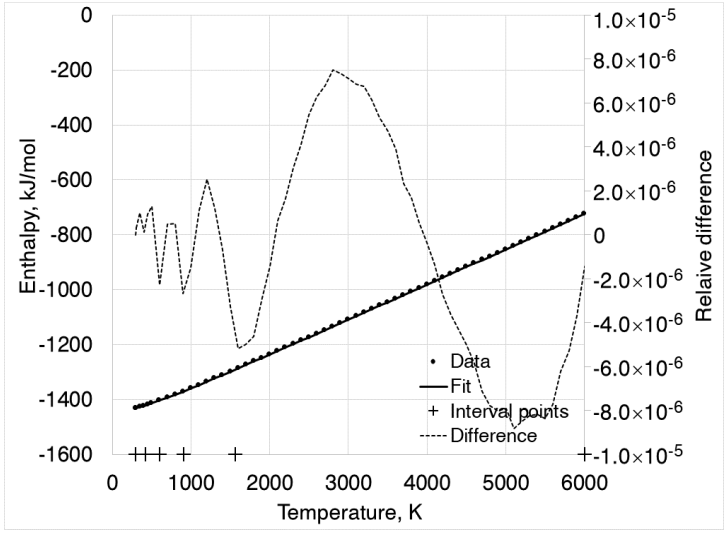


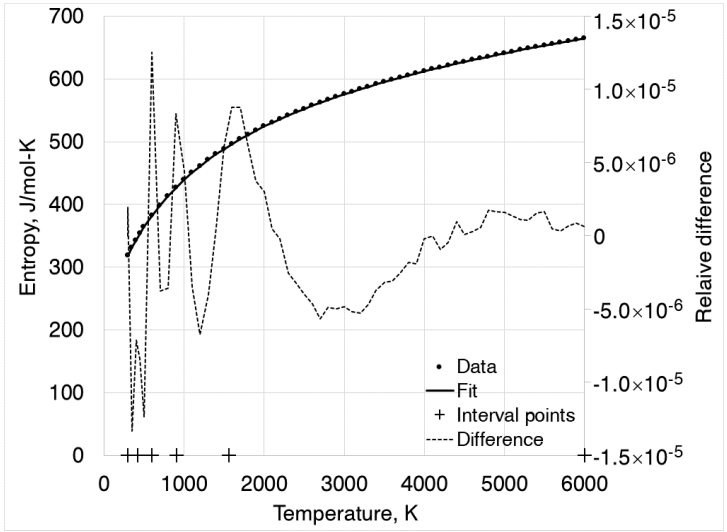


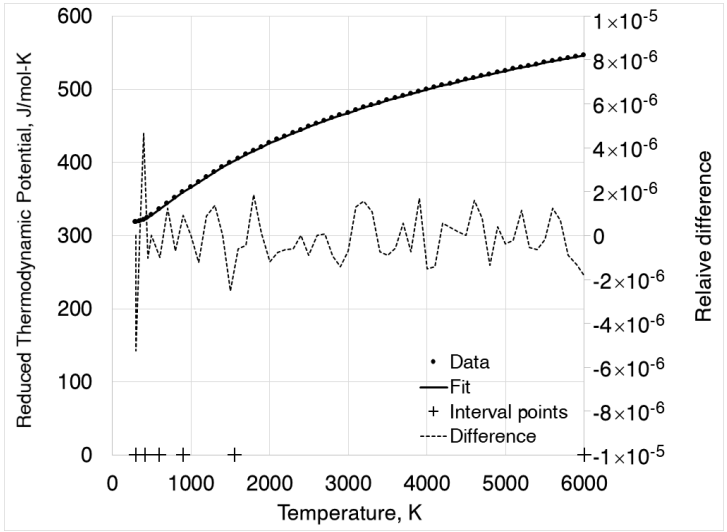
Difluoroborane ($(BF_2)_2$) – B2F4(g)

MW = 97.61 g/mol – $\Delta_f H^o$(298.15 K) = -1431.765 kJ/mol – S^o(298.15 K) = 318.473 J/mol-K – p^o = 0.1 MPa



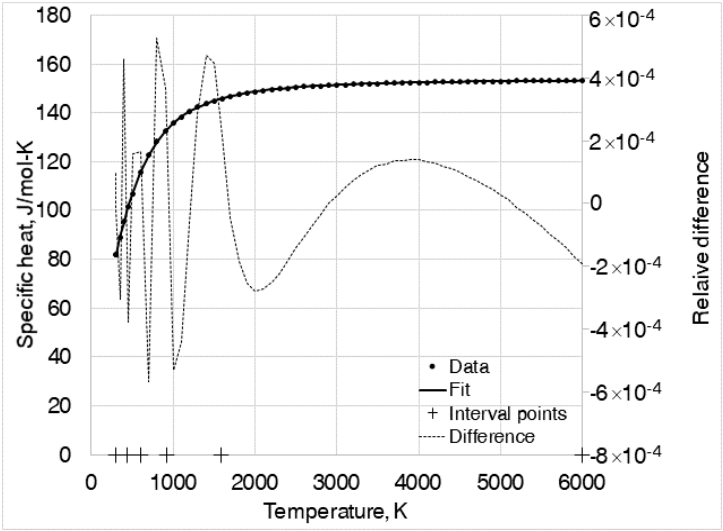


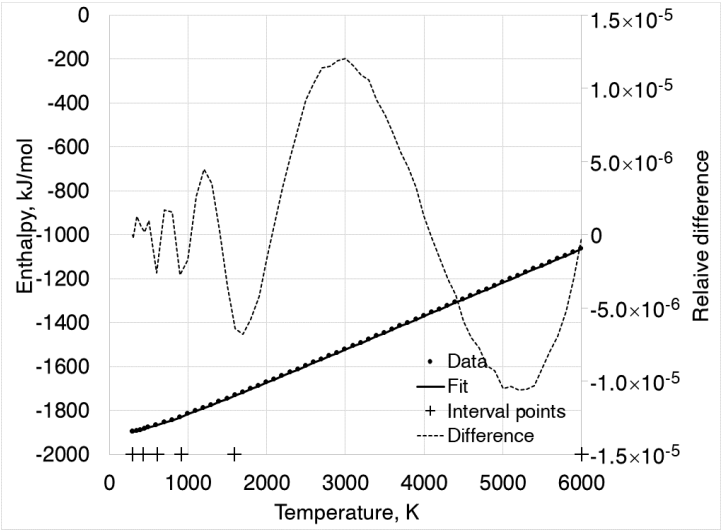


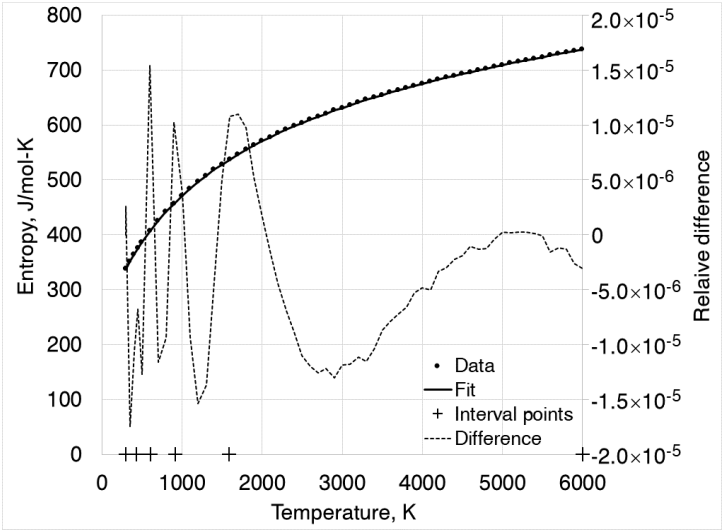


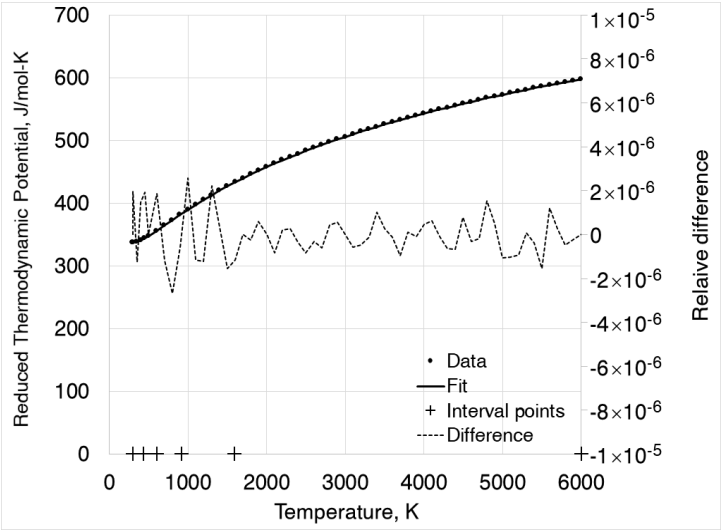
Difluoroborane Oxide ($O(BF_2)_2$) – B2F4O1(g)

MW = 113.61 g/mol – $\Delta_f H^o$(298.15 K) = -1898.072 kJ/mol – S^o(298.15 K) = 337.523 J/mol-K – p^o = 0.1 MPa



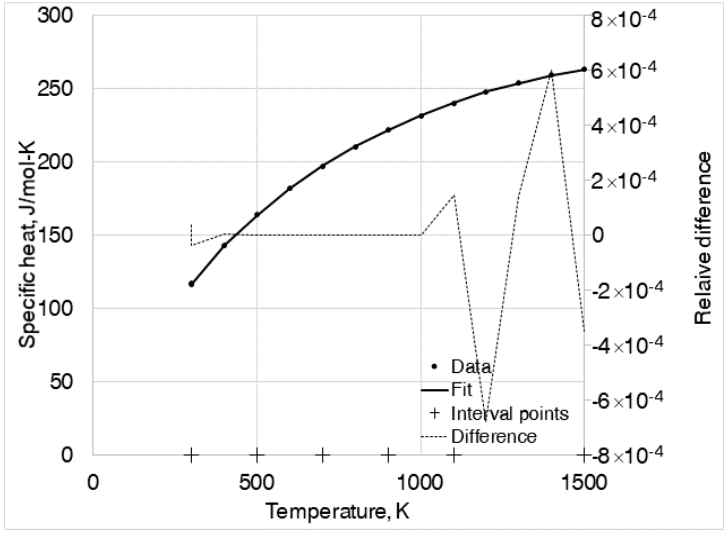


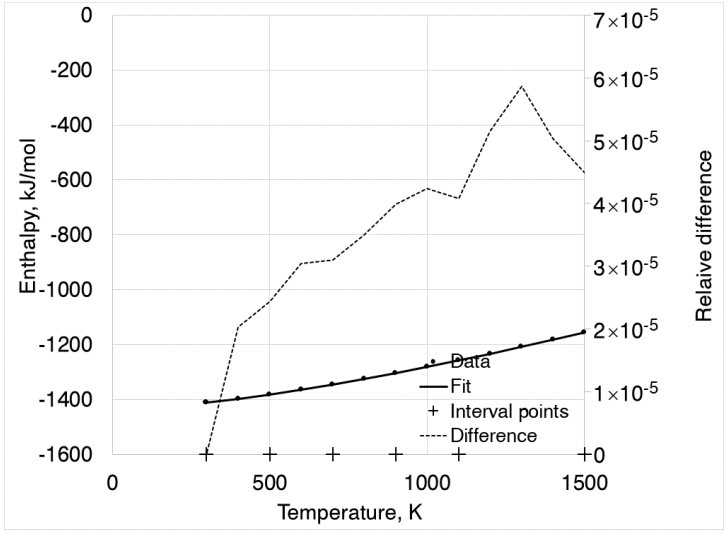


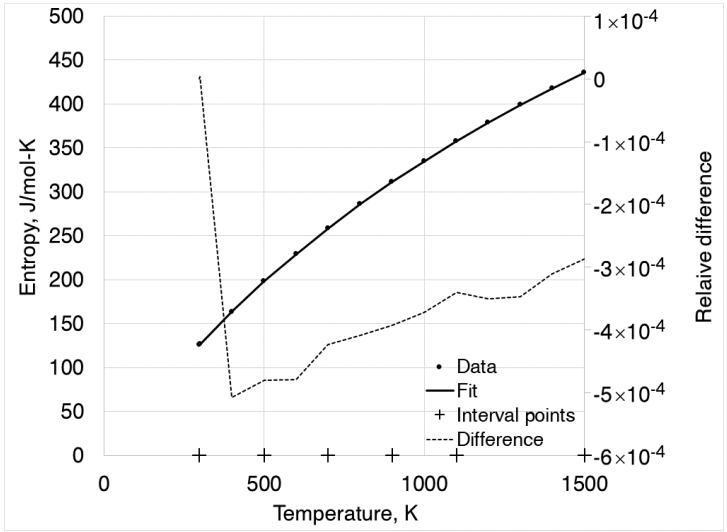


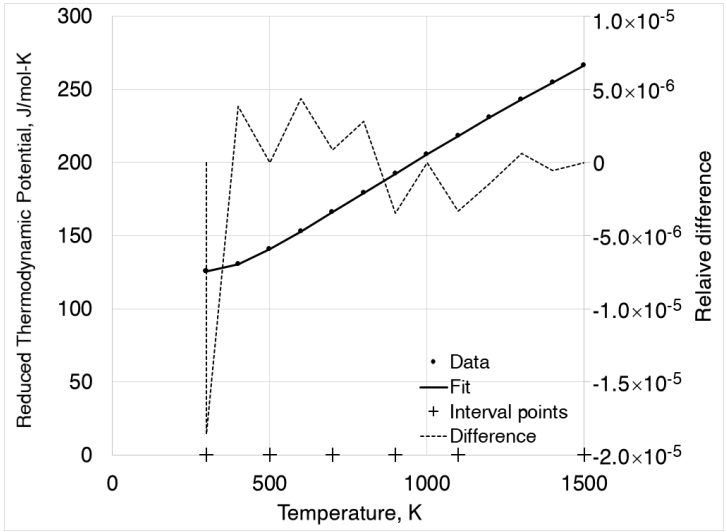
Dihydroxyborane ((B(OH)2)2) – B2H4O4(cr)

MW = 89.65 g/mol – $\Delta_f H^o$(298.15 K) = -1410.426 kJ/mol – S^o(298.15 K) = 125.52 J/mol-K – p^o = 0.1 MPa



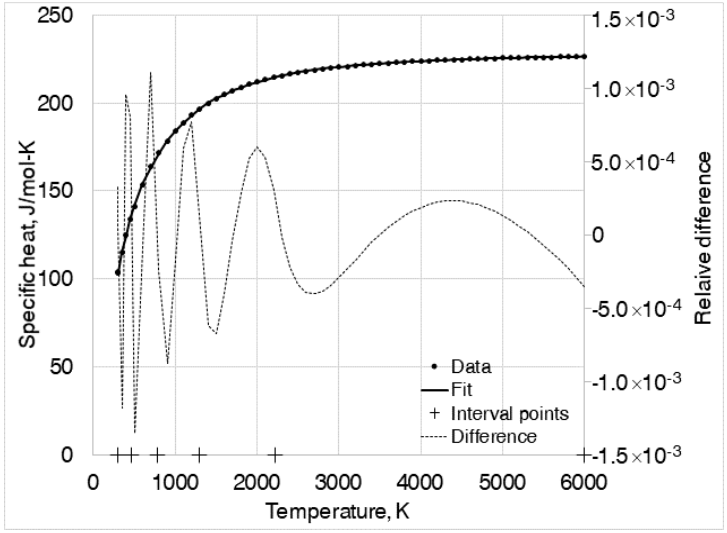


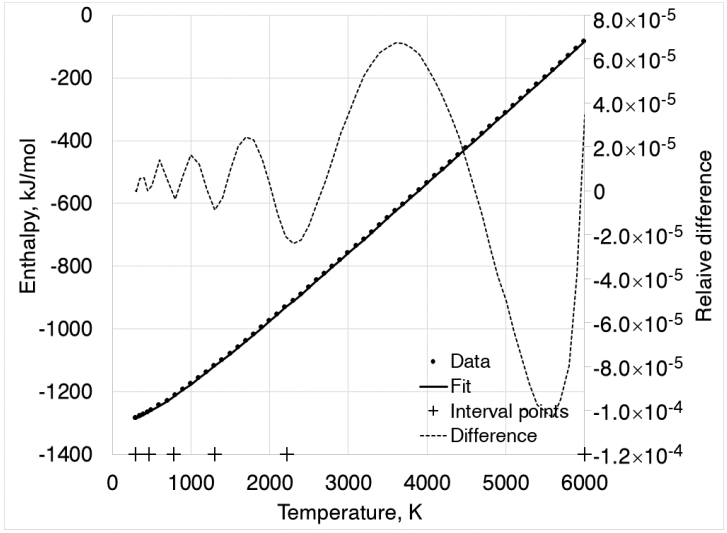


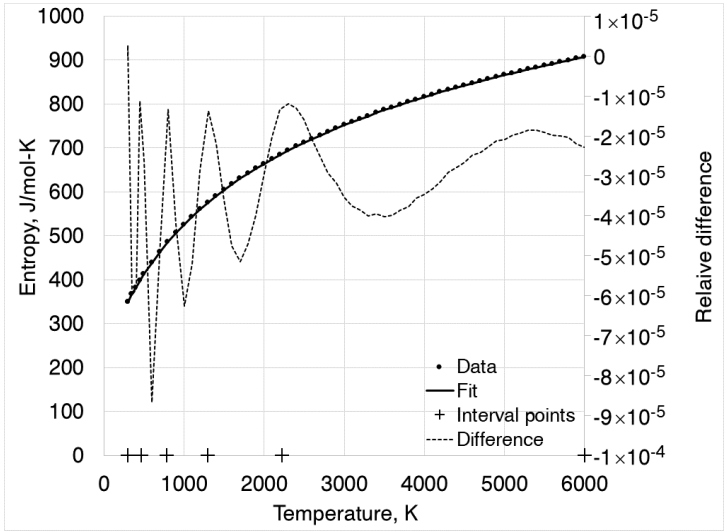


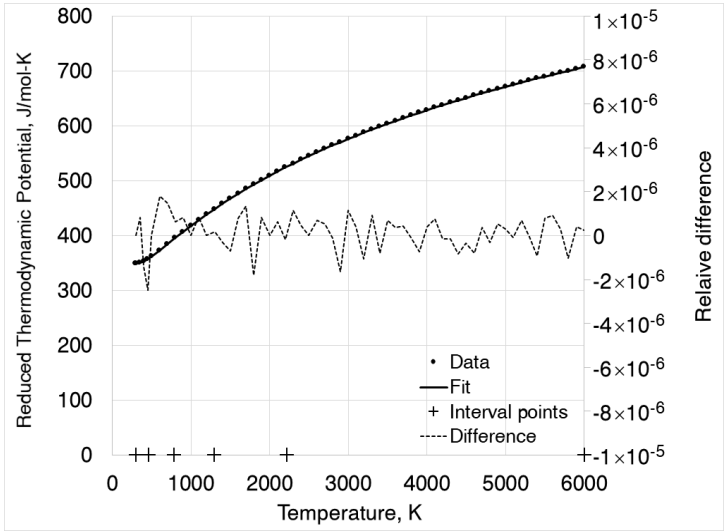
Dihydroxyborane ($(B(OH)_2)_2$) – B2H4O4(g)

MW = 89.65 g/mol – $\Delta_f H^o$(298.15 K) = -1284.488 kJ/mol – S^o(298.15 K) = 348.929 J/mol-K – p^o = 0.1 MPa



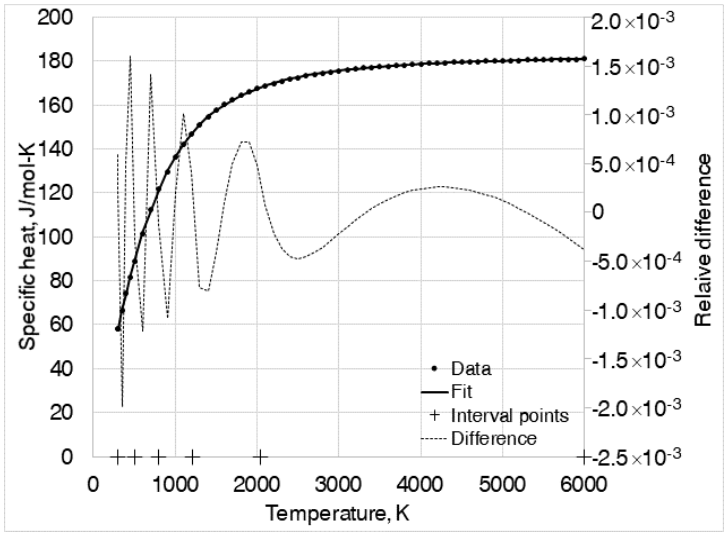


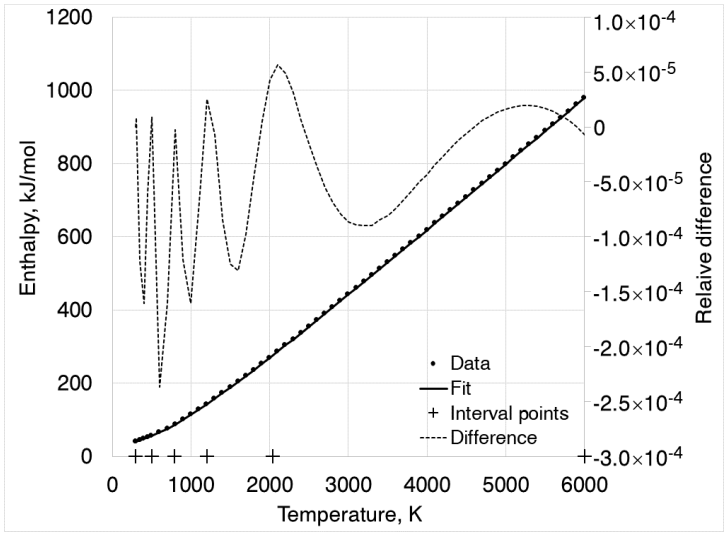


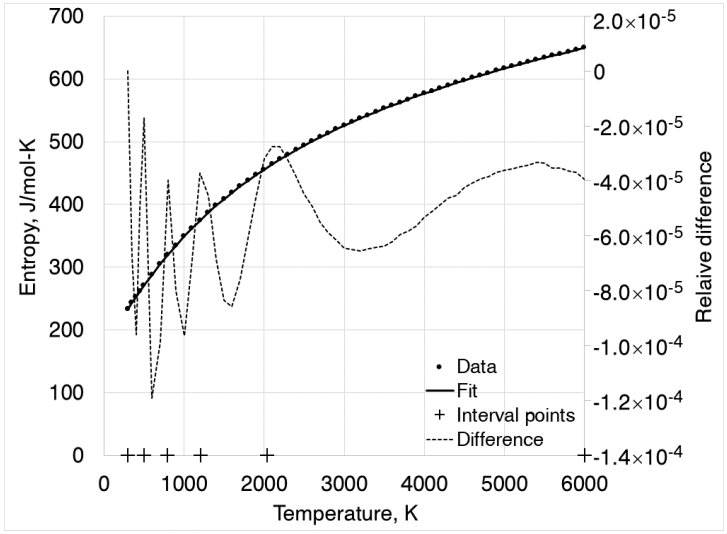


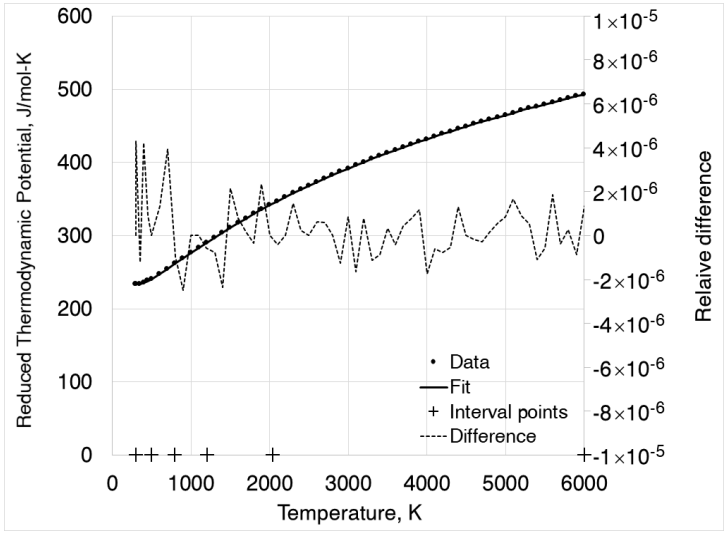
Diborane (B_2H_6) – B2H6(g)

MW = 27.67 g/mol – $\Delta_f H^o$(298.15 K) = 41.003 kJ/mol – S^o(298.15 K) = 233.17 J/mol-K – p^o = 0.1 MPa



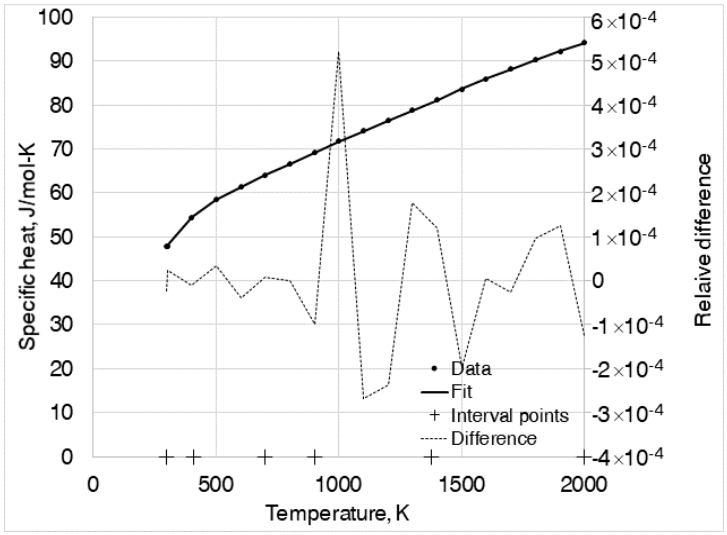


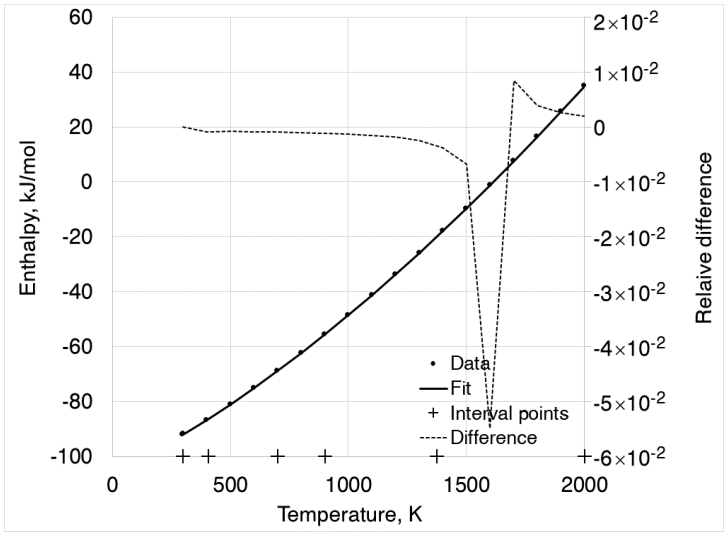


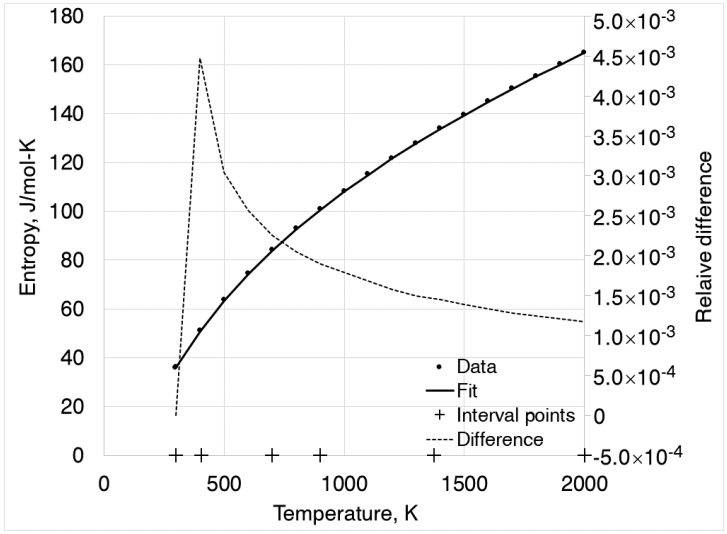


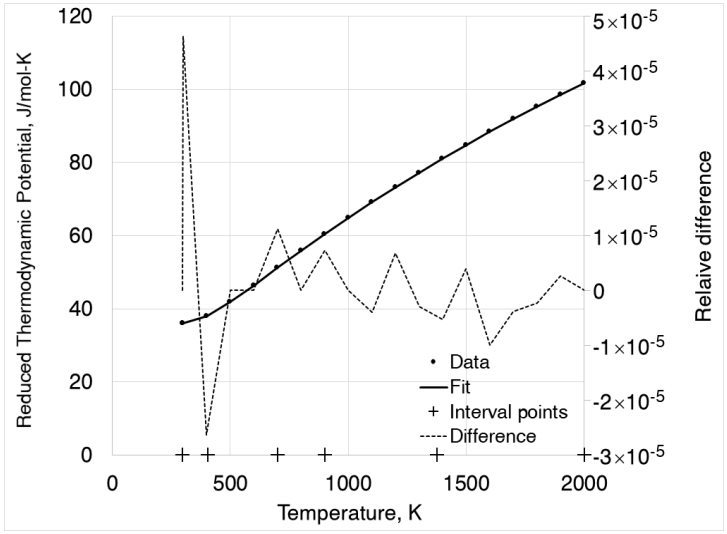
Magnesium Boride (MgB_2) – B2Mg1(cr)

MW = 45.92 g/mol – $\Delta_f H^o$(298.15 K) = -91.964 kJ/mol – S^o(298.15 K) = 35.982 J/mol-K – p^o = 0.1 MPa



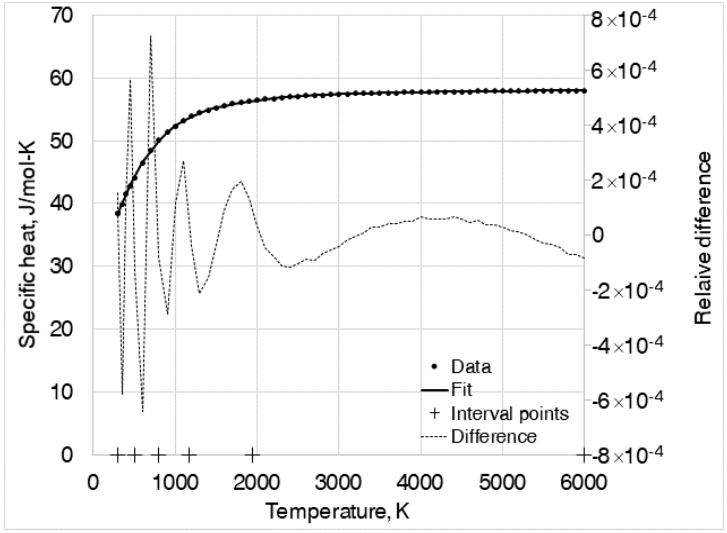


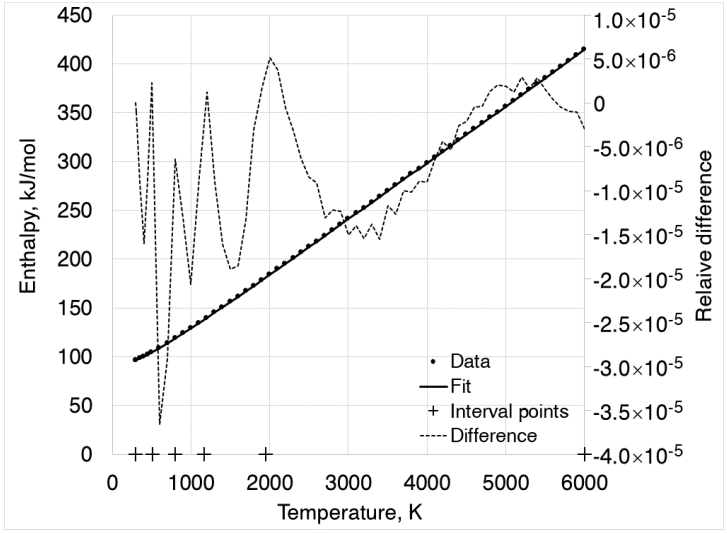


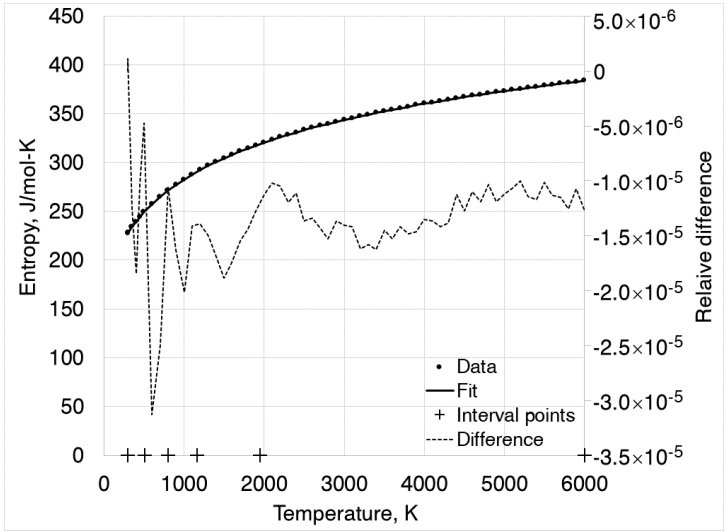


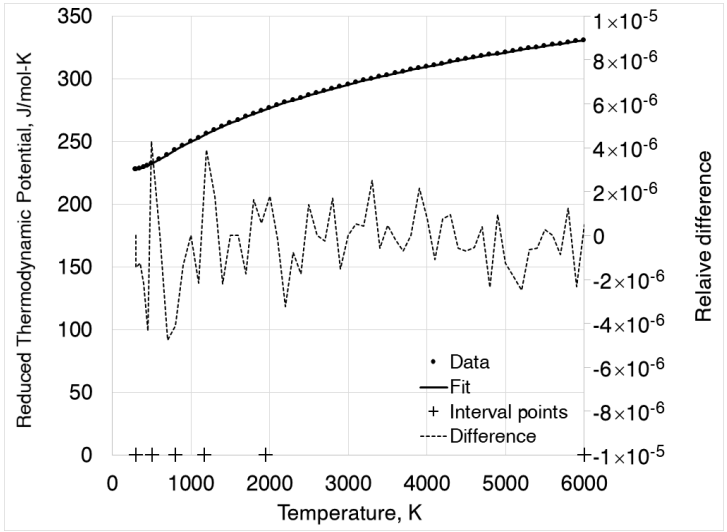
Boron Oxide (B_2O) – B2O1(g)

MW = 37.62 g/mol – $\Delta_f H^o$(298.15 K) = 96.232 kJ/mol – S^o(298.15 K) = 227.744 J/mol-K – p^o = 0.1 MPa



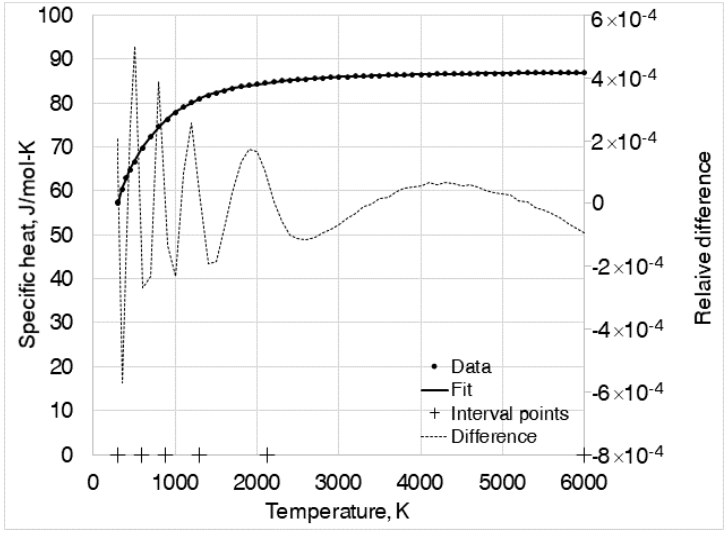


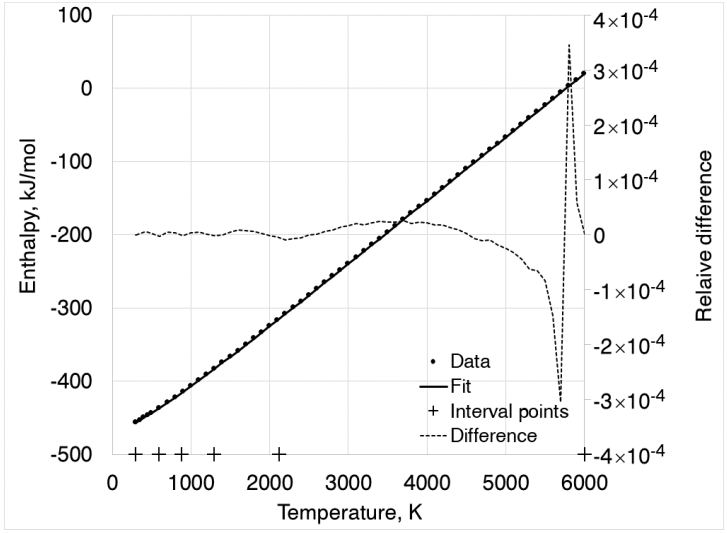


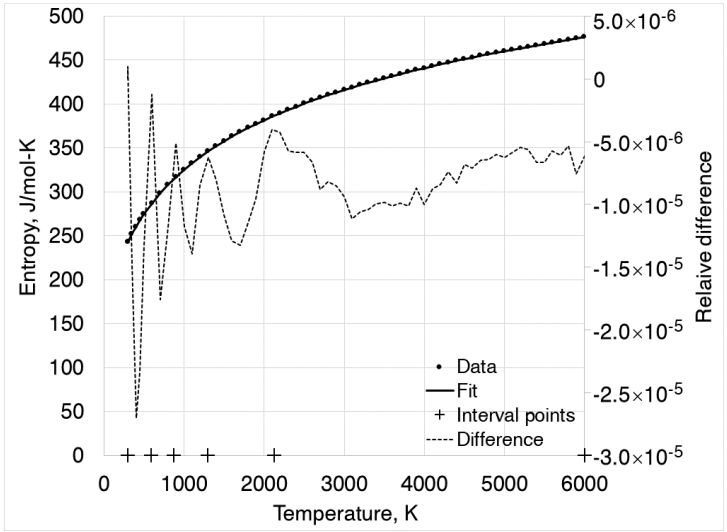


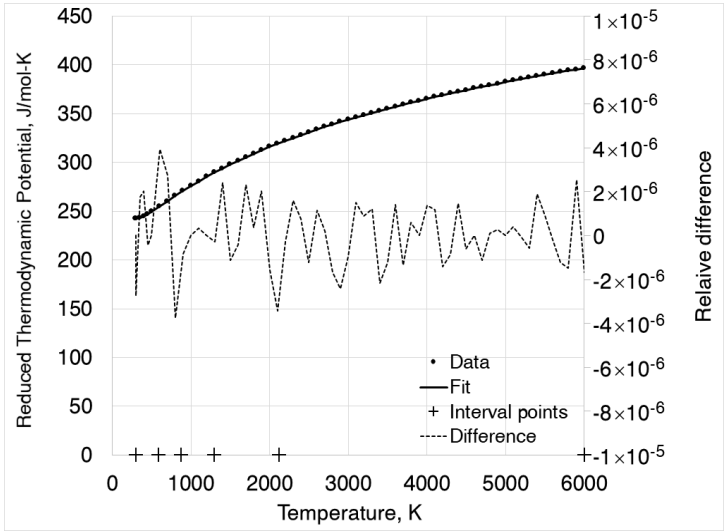
Boron Oxide ($(BO)_2$) – B2O2(g)

MW = 53.62 g/mol – $\Delta_f H^o$(298.15 K) = -456.056 kJ/mol – S^o(298.15 K) = 242.597 J/mol-K – p^o = 0.1 MPa



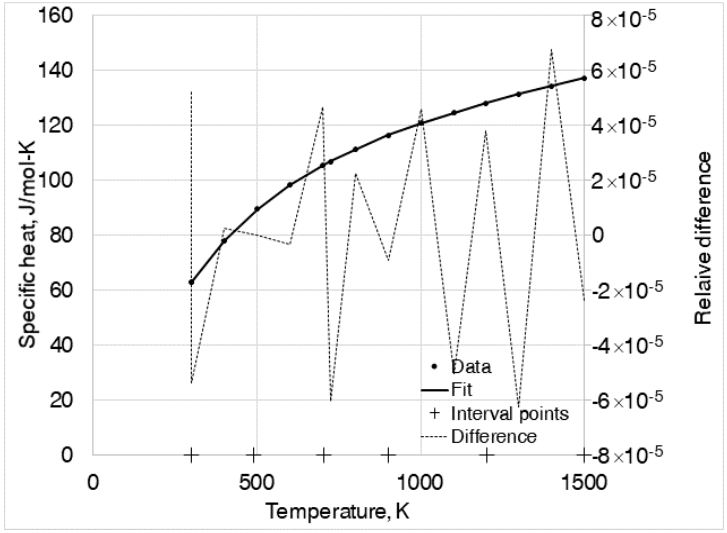


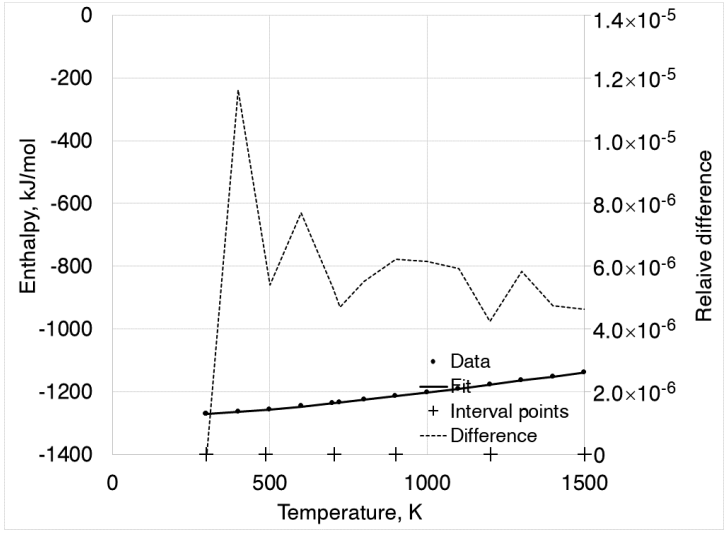


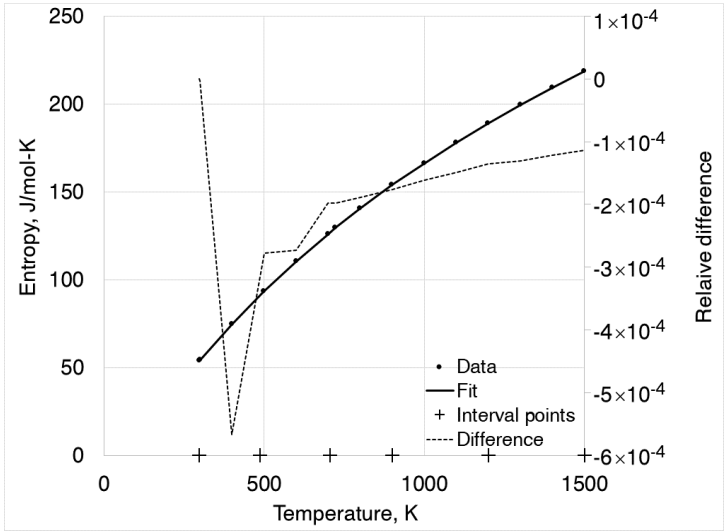


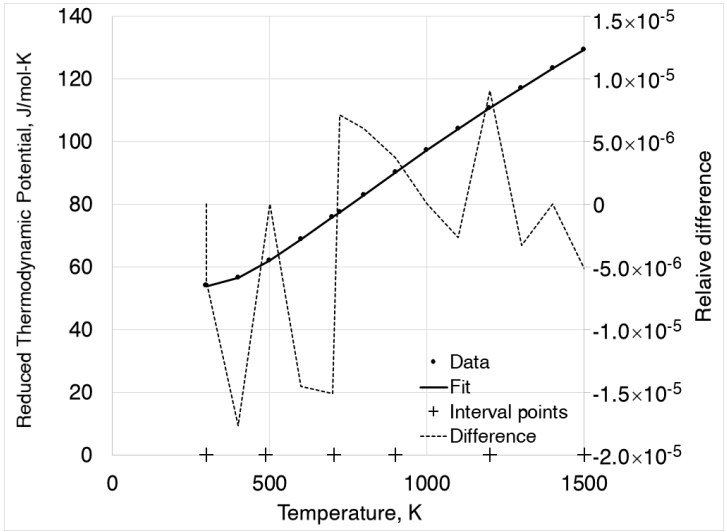
Boron Oxide (B_2O_3) – B2O3(cr)

MW = 69.62 g/mol – $\Delta_f H^o$(298.15 K) = -1271.936 kJ/mol – S^o(298.15 K) = 53.953 J/mol-K – p^o = 0.1 MPa



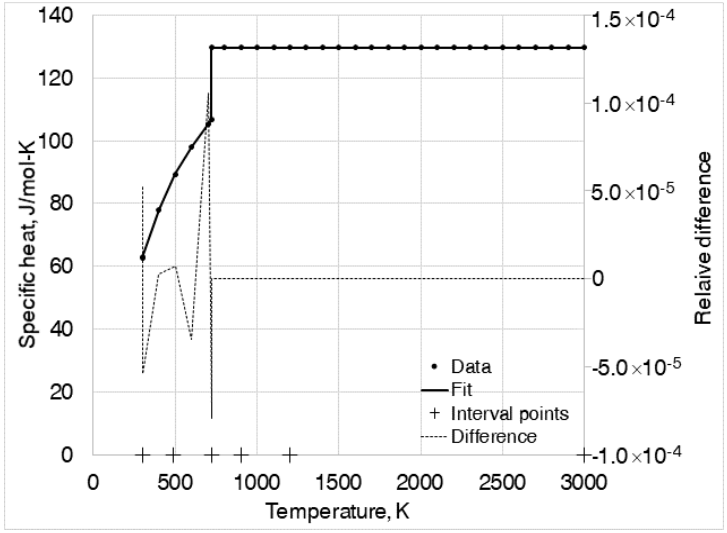


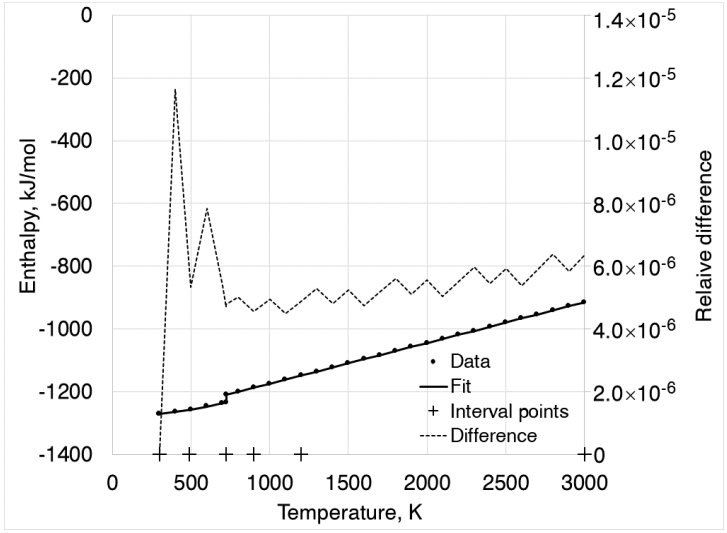


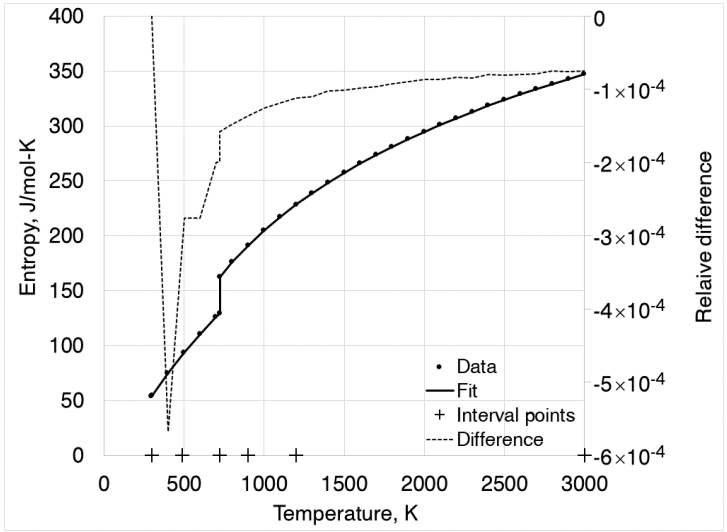


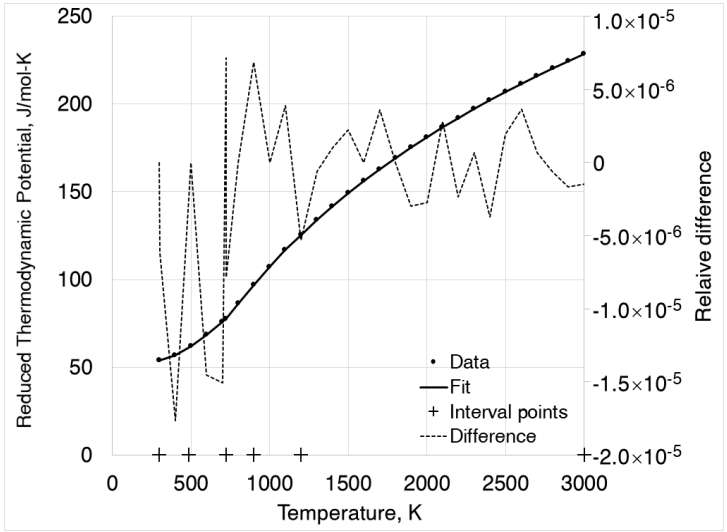
Boron Oxide (B_2O_3) – B2O3(cr,l)

MW = 69.62 g/mol – $\Delta_f H^o$(298.15 K) = -1271.936 kJ/mol – S^o(298.15 K) = 53.953 J/mol-K – p^o = 0.1 MPa



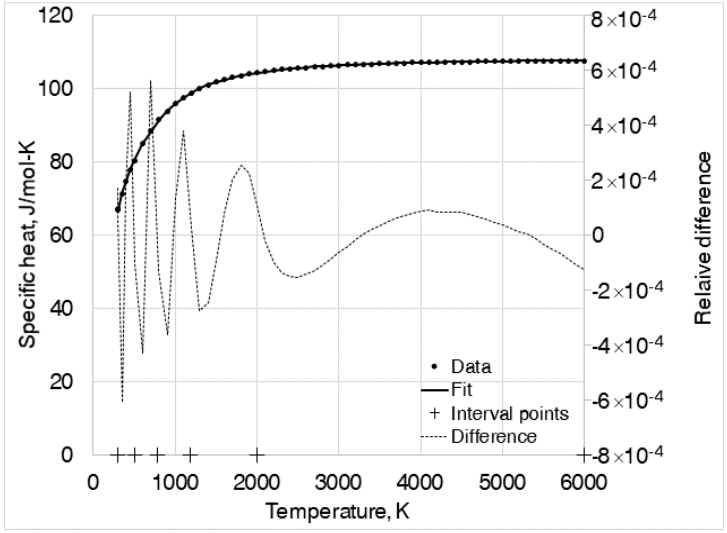


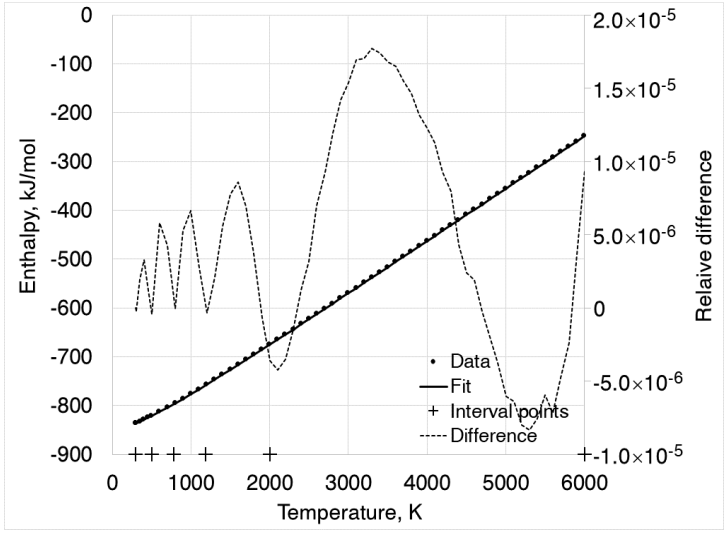


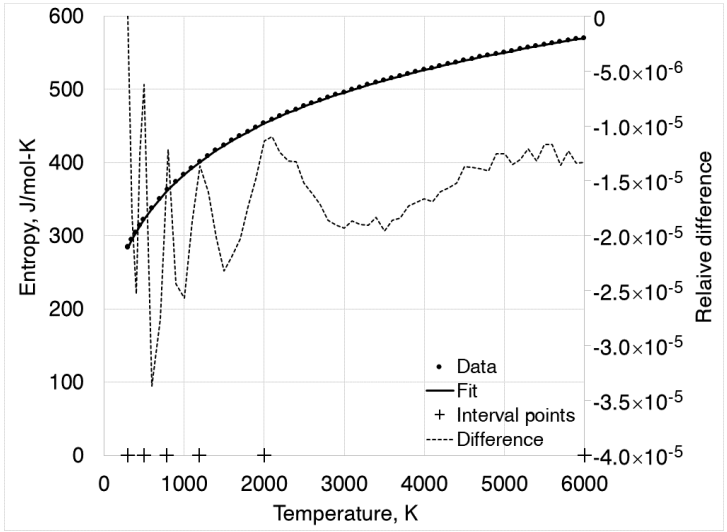


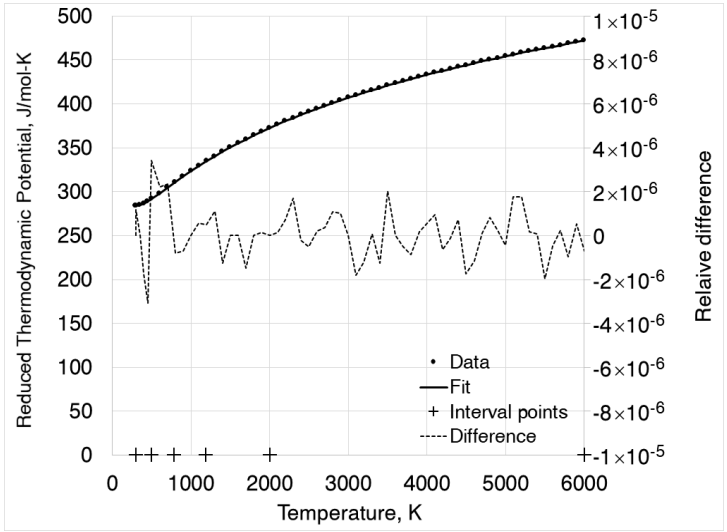
Boron Oxide (B_2O_3) – B2O3(g)

MW = 69.62 g/mol – $\Delta_f H^o$(298.15 K) = -835.963 kJ/mol – S^o(298.15 K) = 283.768 J/mol-K – p^o = 0.1 MPa



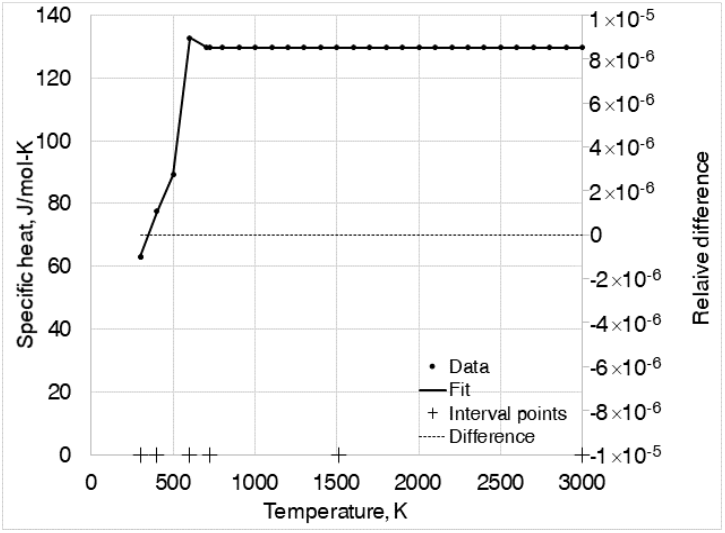


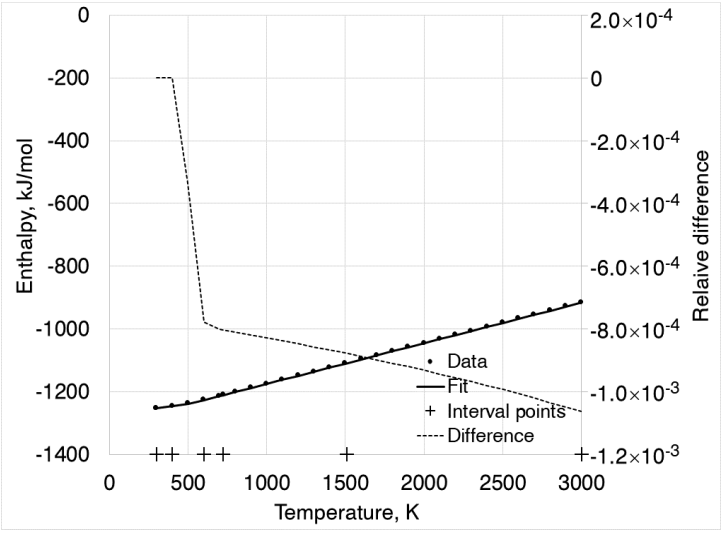


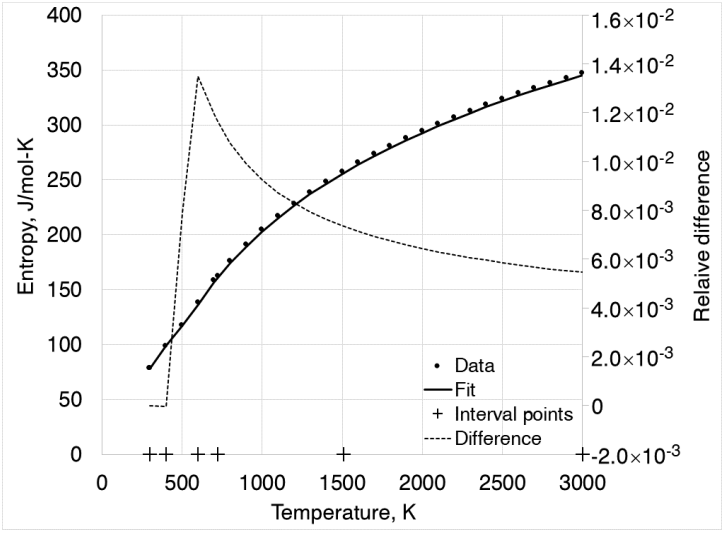


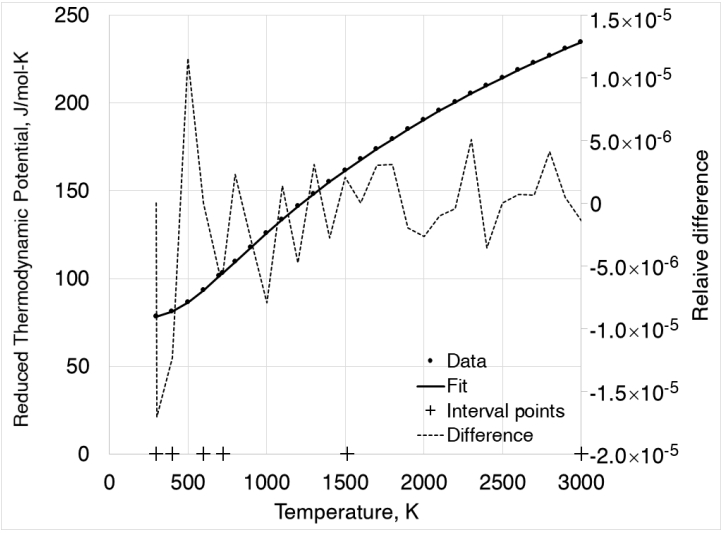
Boron Oxide (B_2O_3) – B2O3(l)

MW = 69.62 g/mol – $\Delta_f H^o$(298.15 K) = -1253.359 kJ/mol – S^o(298.15 K) = 78.443 J/mol-K – p^o = 0.1 MPa



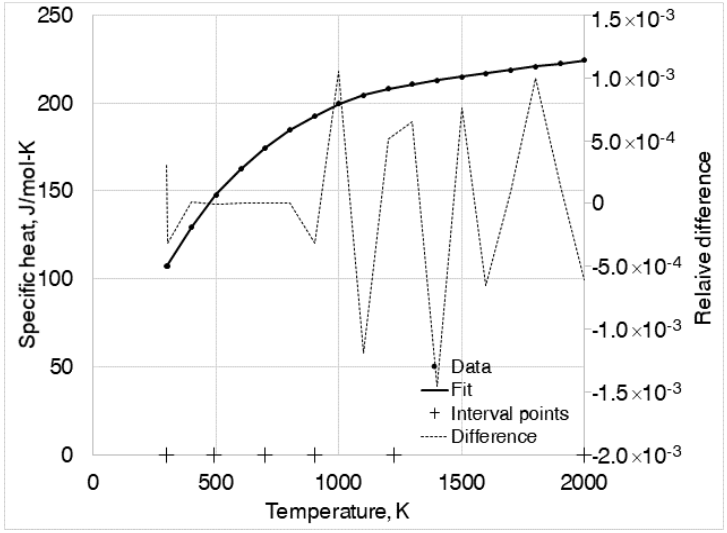


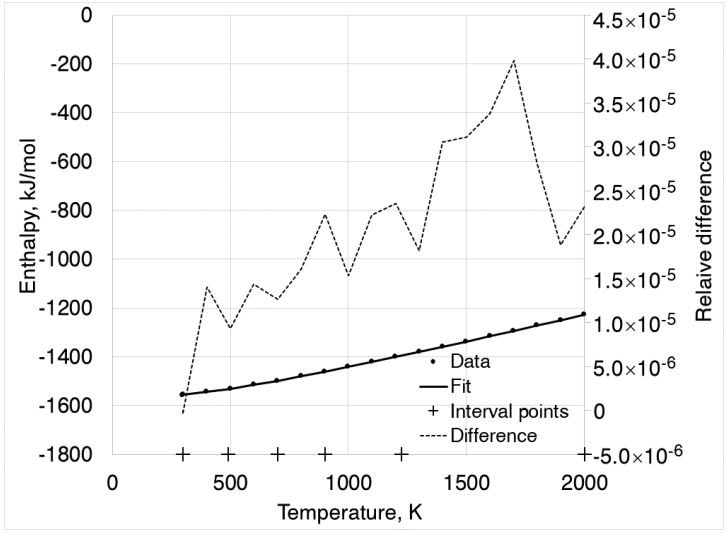


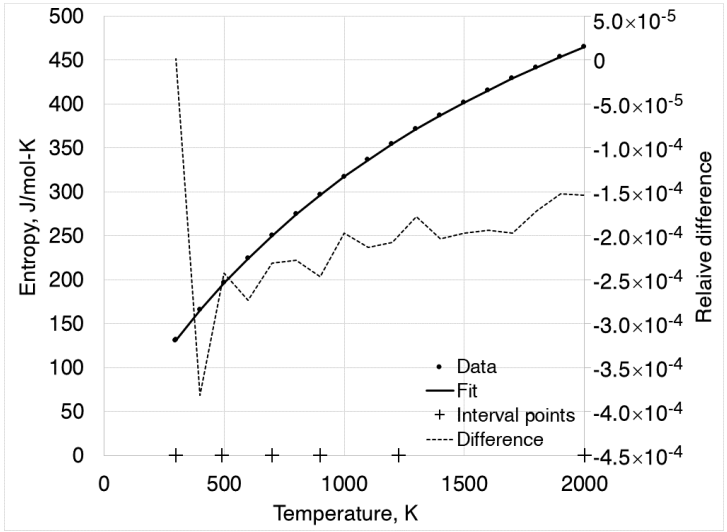


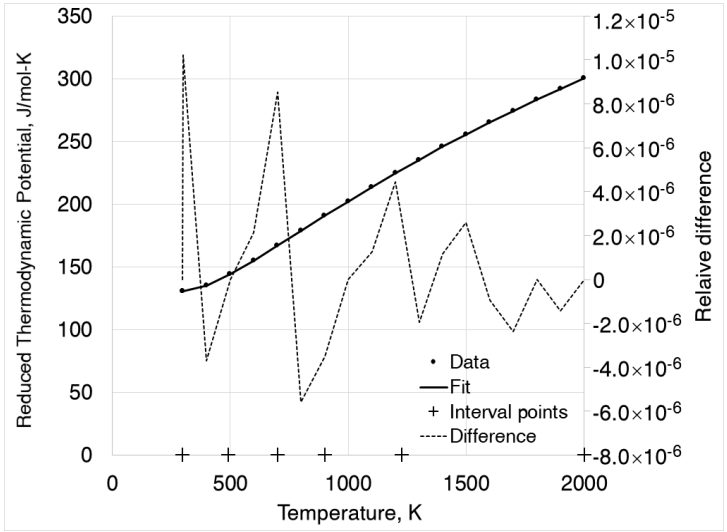
Lead Borate (PbB_2O_4) – B2O4Pb1(cr)

MW = 292.8 g/mol – $\Delta_f H^o$(298.15 K) = -1556.448 kJ/mol – S^o(298.15 K) = 130.541 J/mol-K – p^o = 0.1 MPa



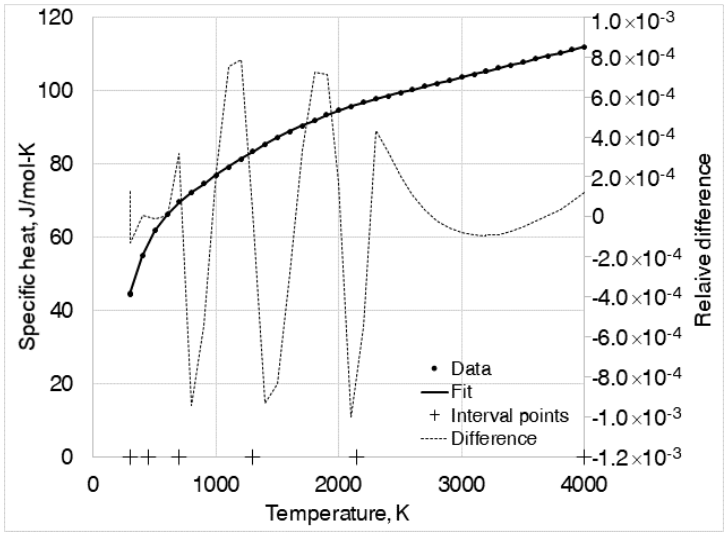


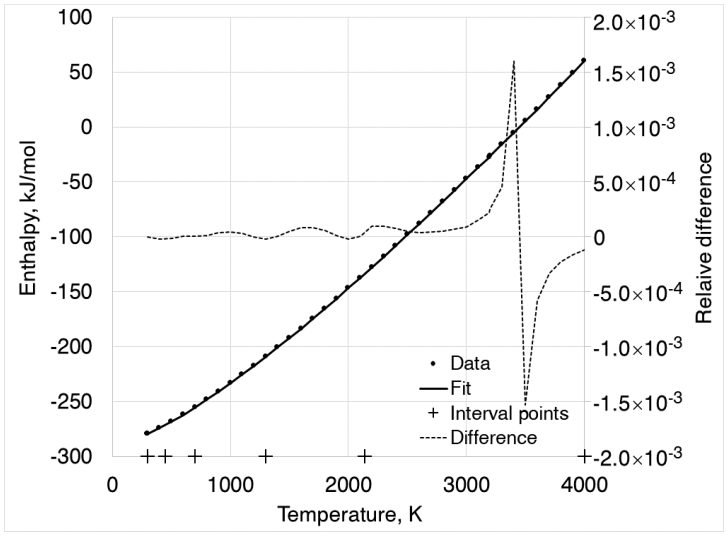


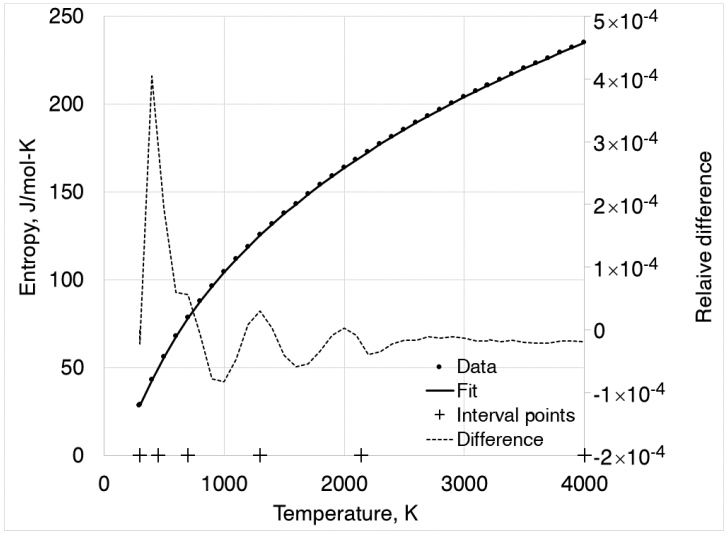


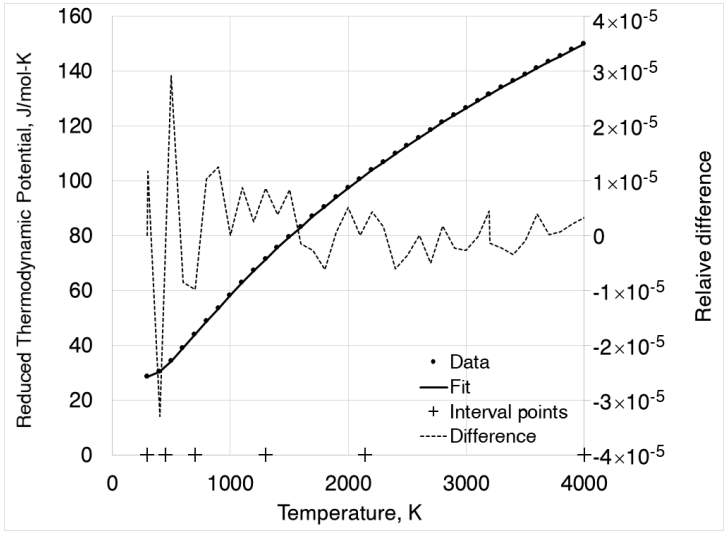
Titanium Boride (TiB_2) – B2Ti1(cr)

MW = 69.5 g/mol – $\Delta_f H^o$(298.15 K) = -279.491 kJ/mol – S^o(298.15 K) = 28.485 J/mol-K – p^o = 0.1 MPa



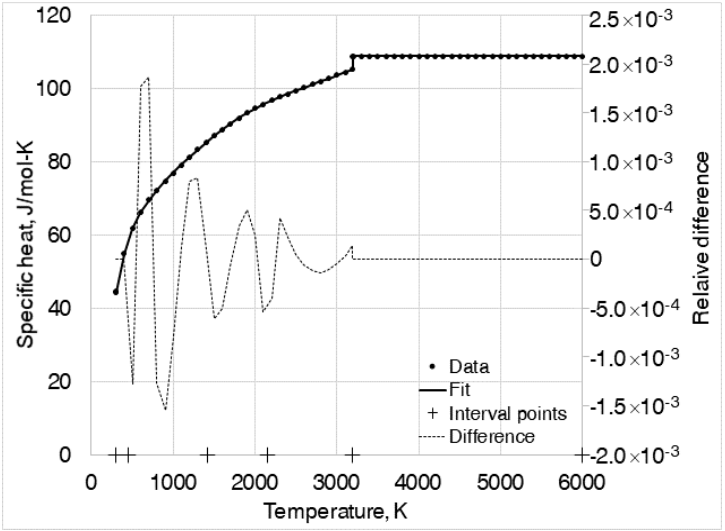


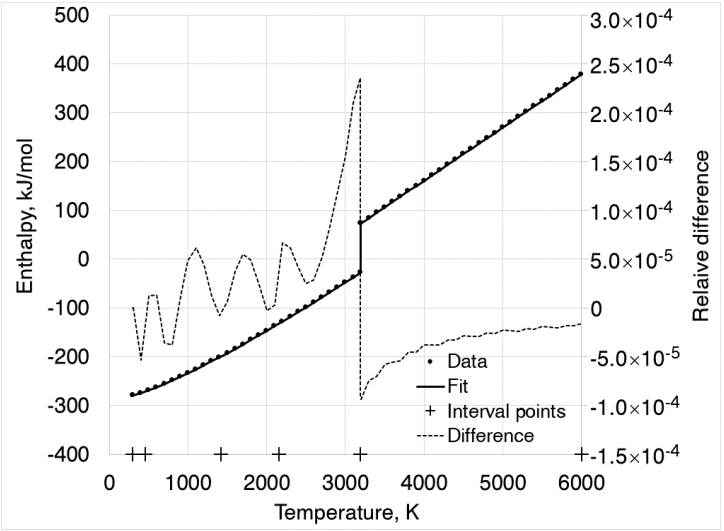


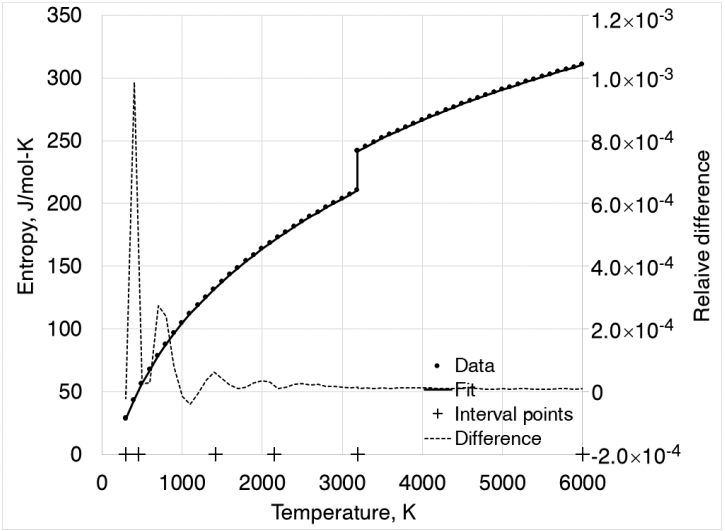


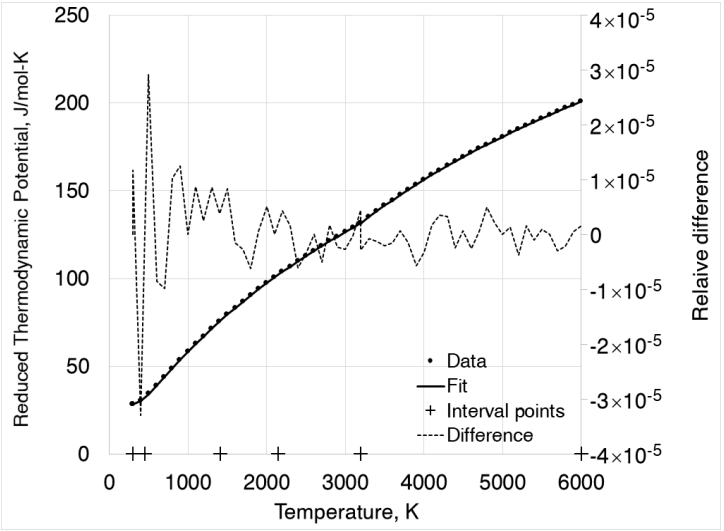
Titanium Boride (TiB_2) – B2Ti1(cr,l)

MW = 69.5 g/mol – $\Delta_f H^o$(298.15 K) = -279.491 kJ/mol – S^o(298.15 K) = 28.485 J/mol-K – p^o = 0.1 MPa



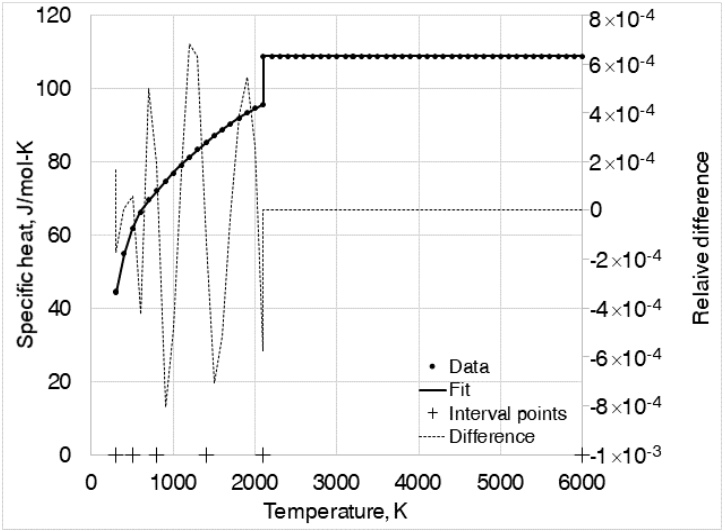


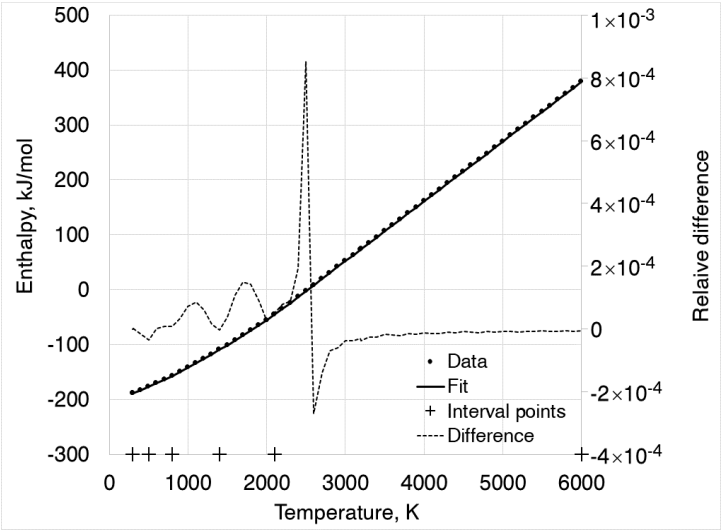


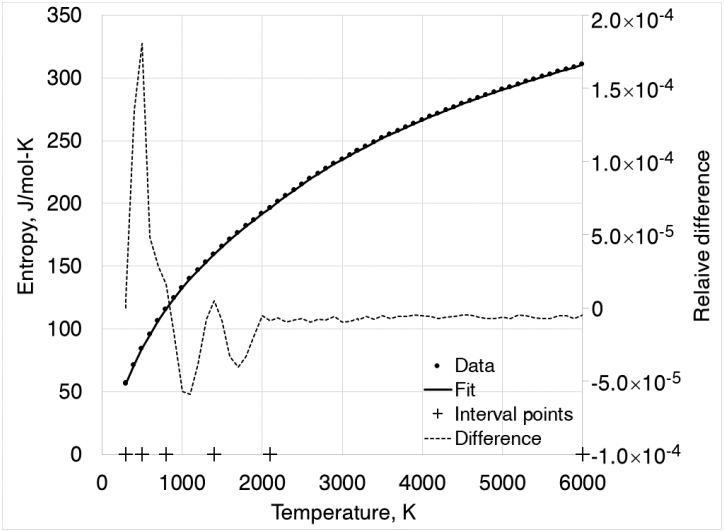


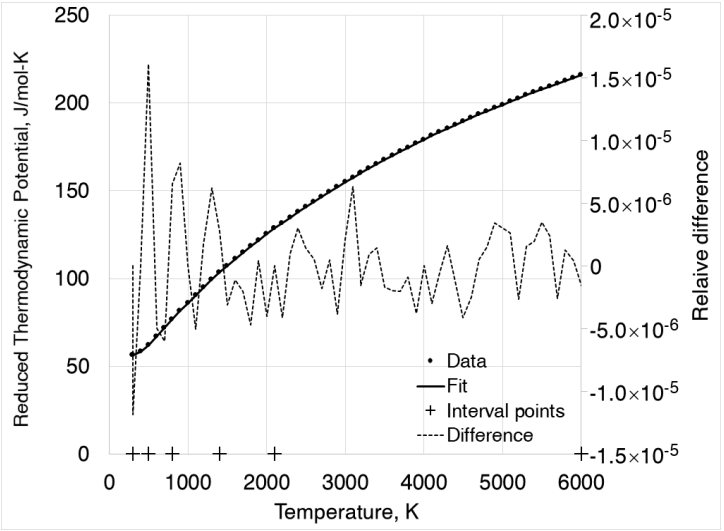
Titanium Boride (TiB_2) – B2Ti1(l)

MW = 69.5 g/mol – $\Delta_f H^o$(298.15 K) = -188.029 kJ/mol – S^o(298.15 K) = 56.365 J/mol-K – p^o = 0.1 MPa



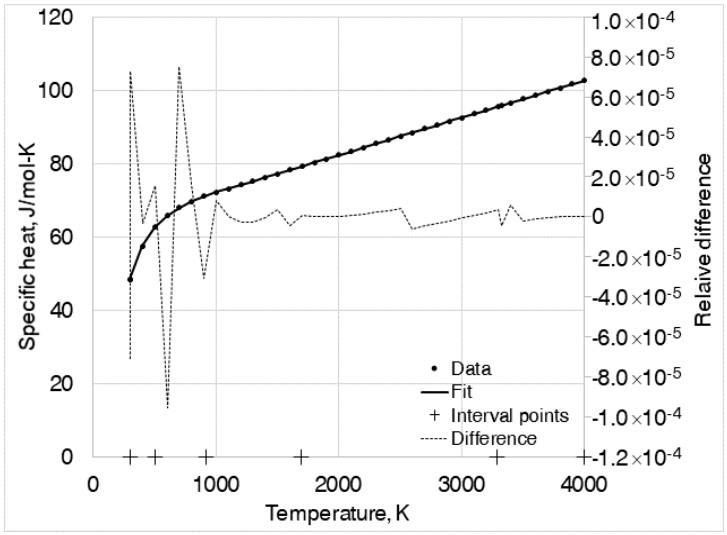


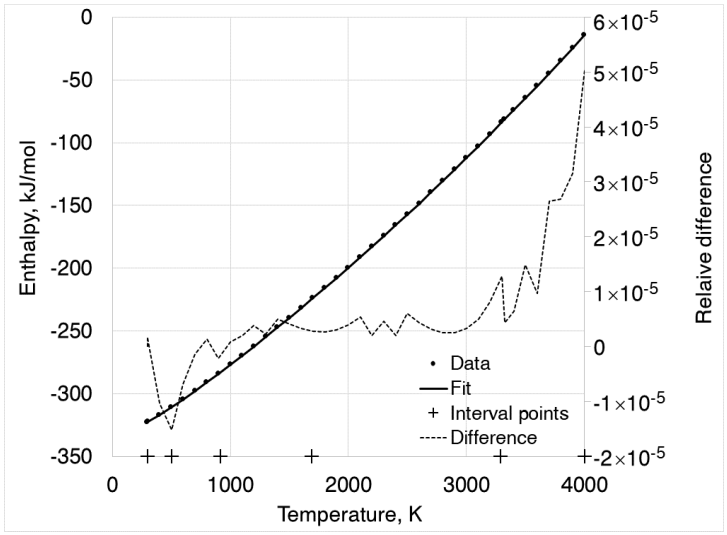


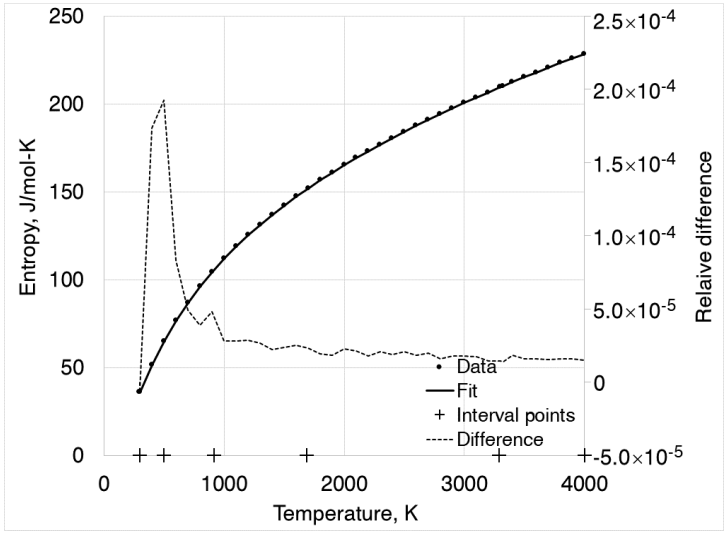


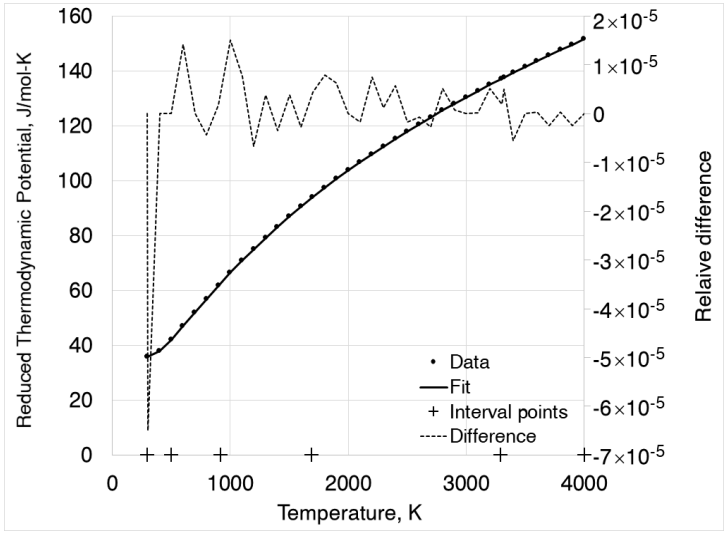
Zirconium Boride (ZrB_2) – B2Zr1(cr)

MW = 112.84 g/mol – $\Delta_f H^o$(298.15 K) = -322.586 kJ/mol – S^o(298.15 K) = 35.941 J/mol-K – p^o = 0.1 MPa



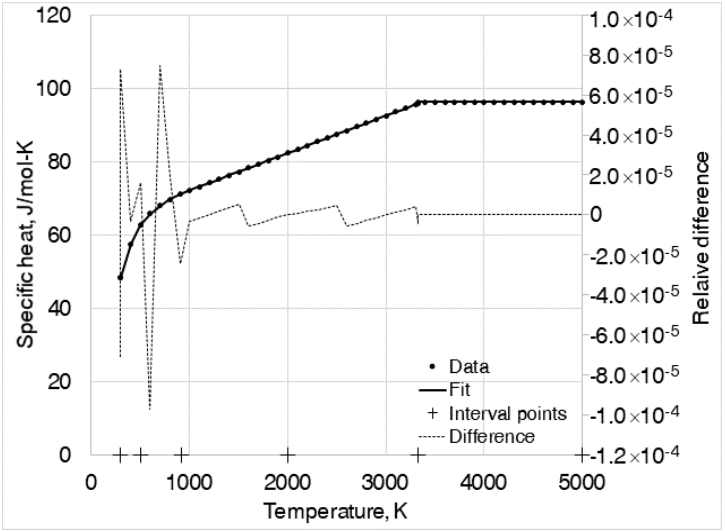


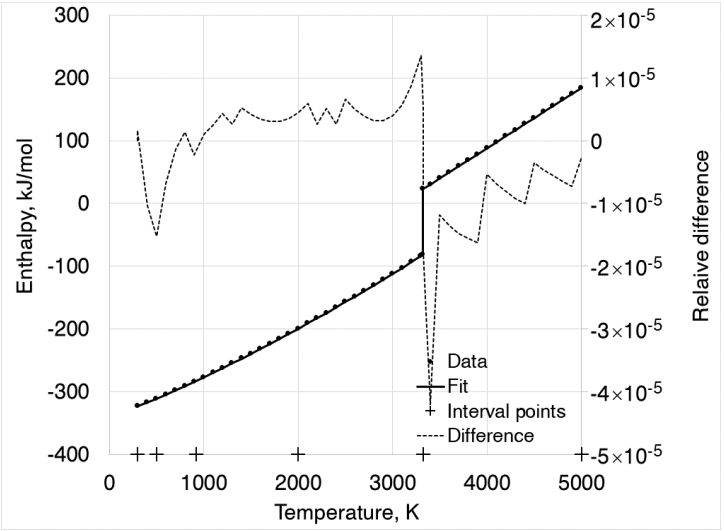


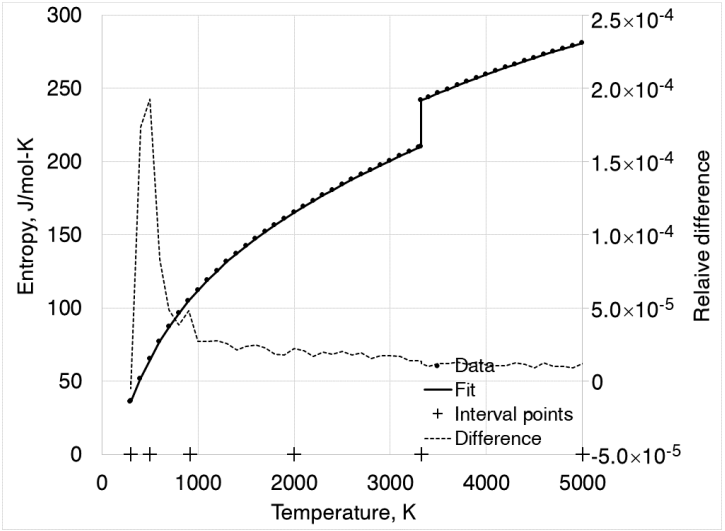


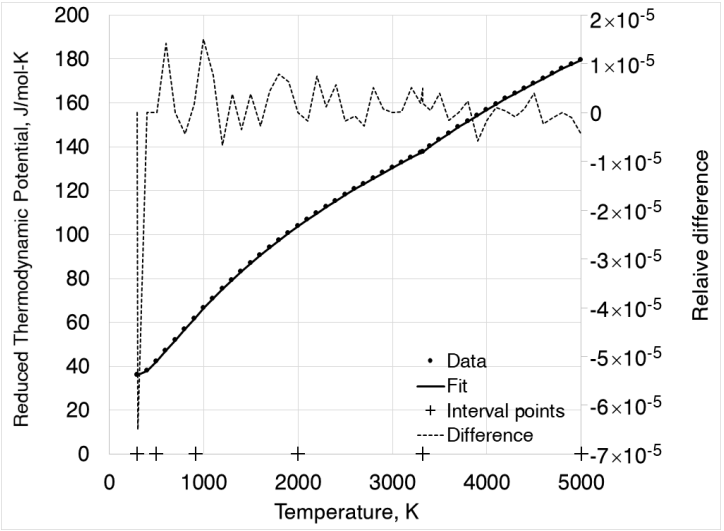
Zirconium Boride (ZrB_2) – B2Zr1(cr,l)

MW = 112.84 g/mol – $\Delta_f H^o$(298.15 K) = -322.586 kJ/mol – S^o(298.15 K) = 35.941 J/mol-K – p^o = 0.1 MPa



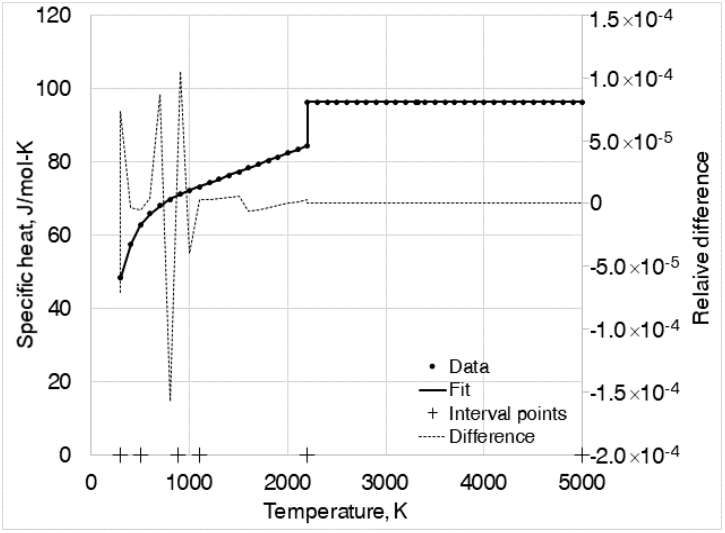


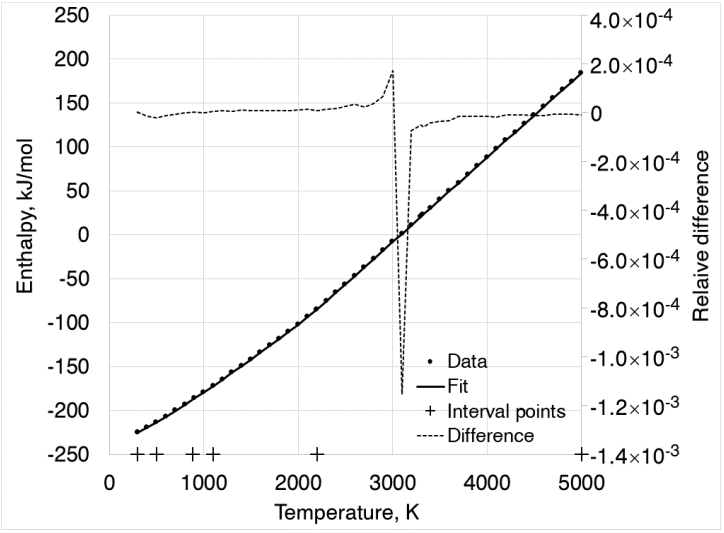


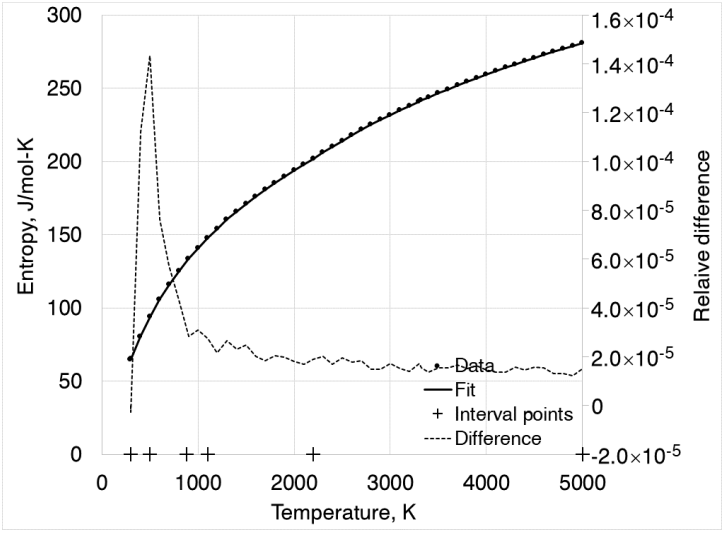


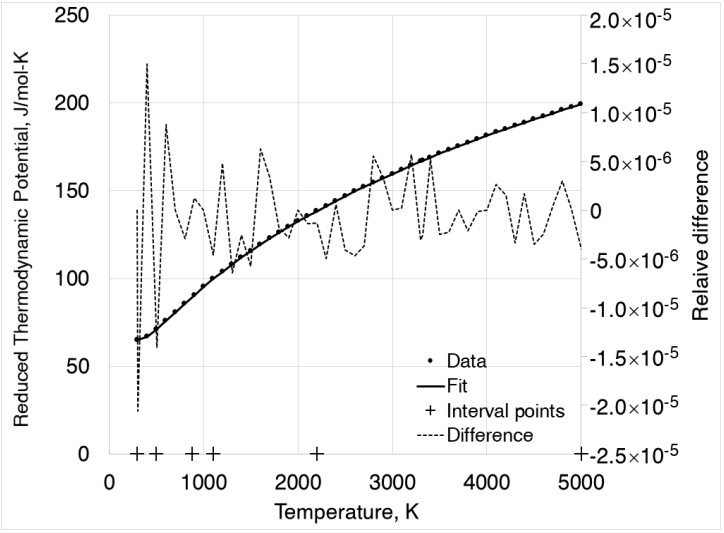
Zirconium Boride (ZrB_2) – B2Zr1(l)

MW = 112.84 g/mol – $\Delta_f H^o$(298.15 K) = -224.832 kJ/mol – S^o(298.15 K) = 64.742 J/mol-K – p^o = 0.1 MPa



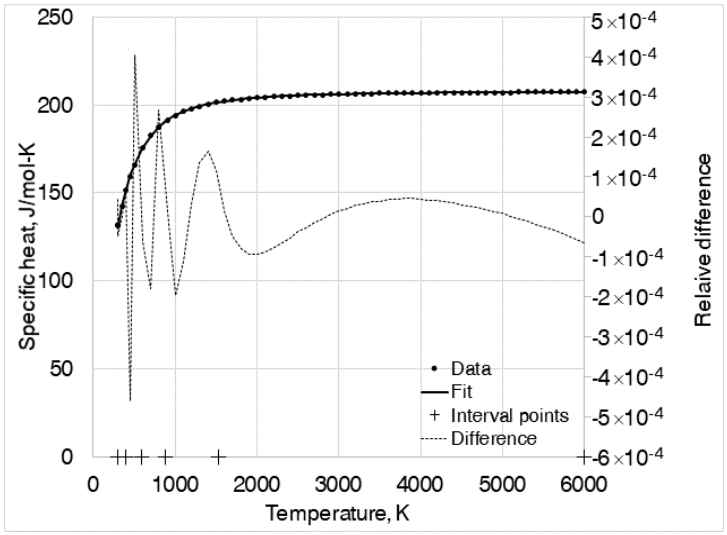


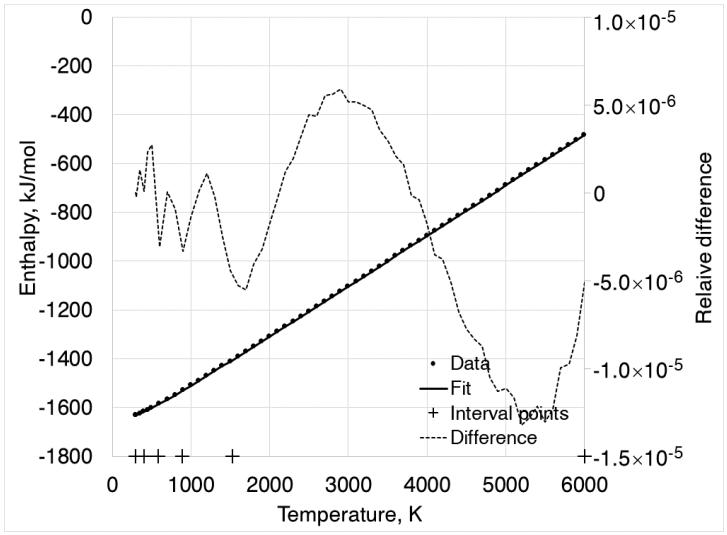


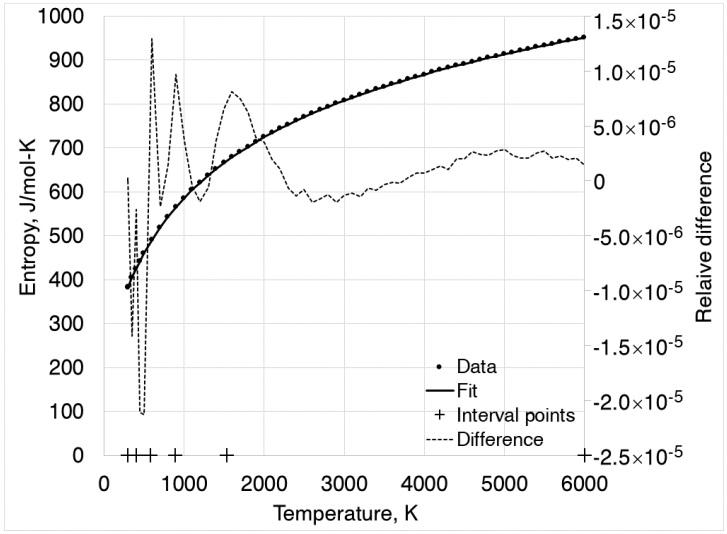


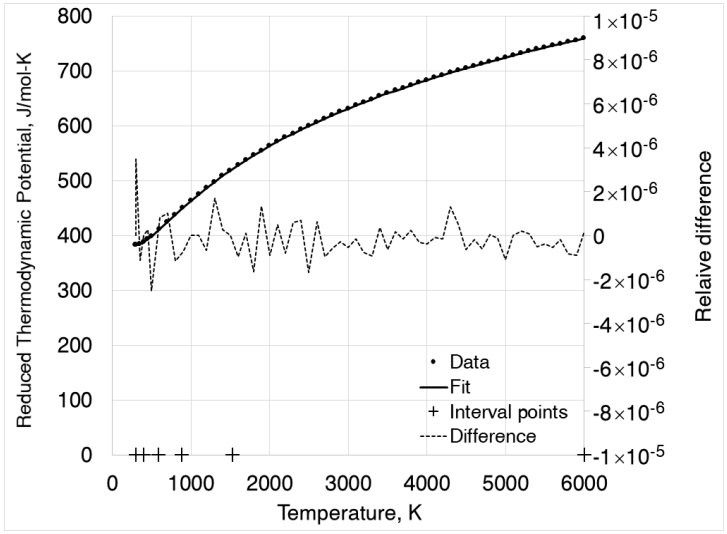
Trichloroboroxin ($B_3O_3Cl_3$) – B3Cl3O3(g)

MW = 186.79 g/mol – $\Delta_f H^o$(298.15 K) = -1631.76 kJ/mol – S^o(298.15 K) = 382.378 J/mol-K – p^o = 0.1 MPa



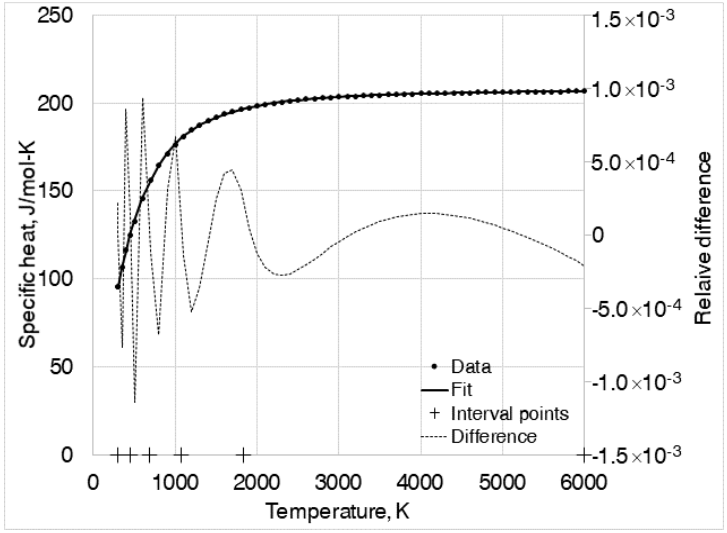


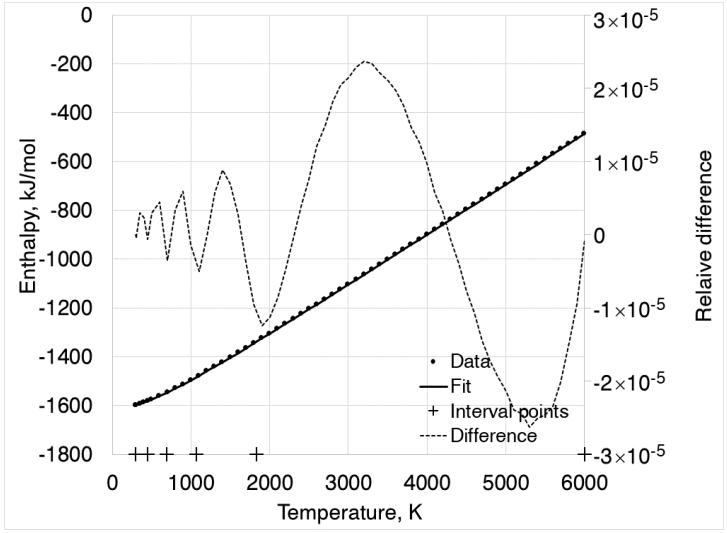


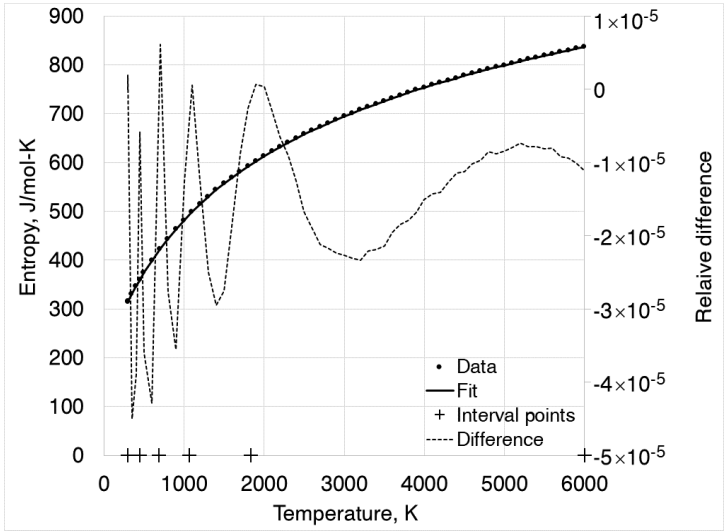


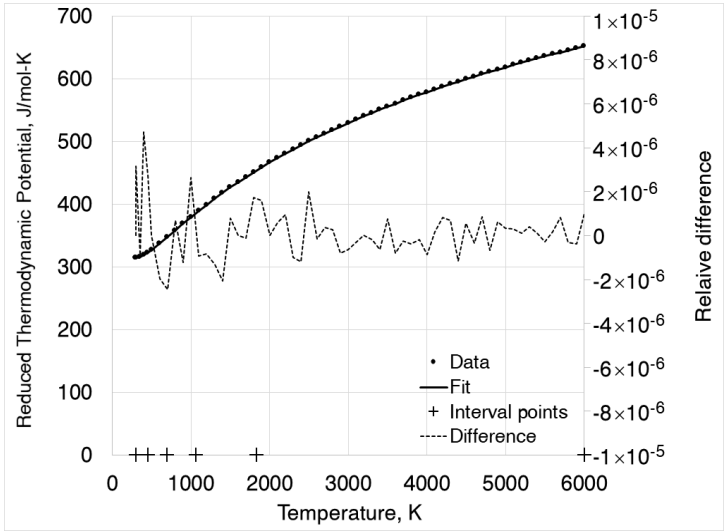
Fluoroboroxin ($B_3H_2O_3F$) – B3F1H2O3(g)

MW = 101.44 g/mol – $\Delta_f H^o$(298.15 K) = -1598.288 kJ/mol – S^o(298.15 K) = 314.841 J/mol-K – p^o = 0.1 MPa



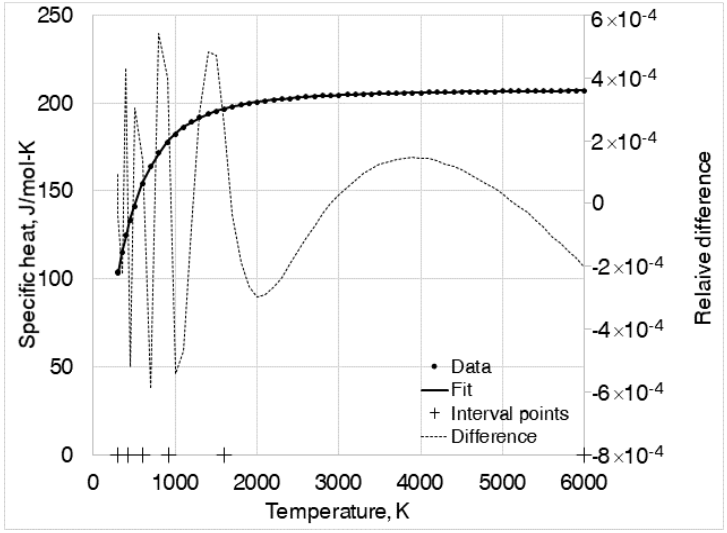


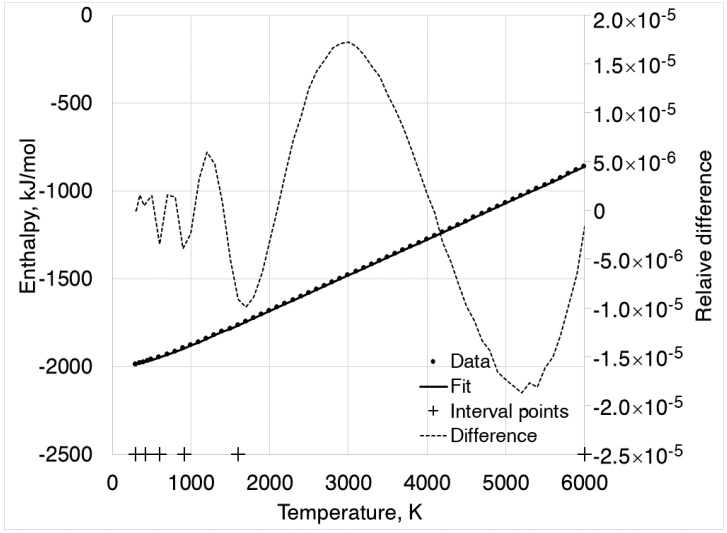


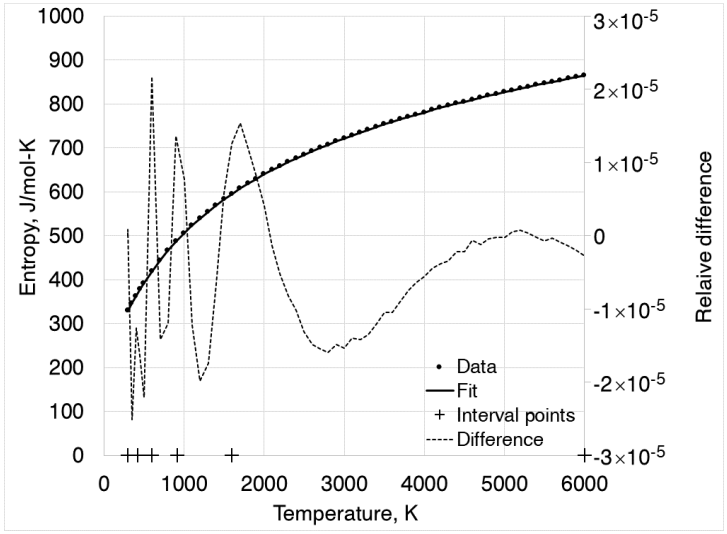


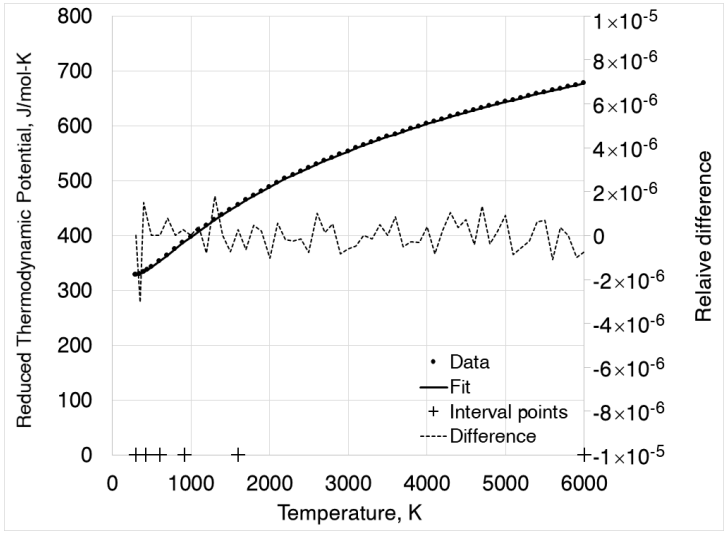
Difluoroboroxin ($B_3HO_3F_2$) – B3F2H1O3(g)

MW = 119.43 g/mol – $\Delta_f H^o$(298.15 K) = -1987.4 kJ/mol – S^o(298.15 K) = 328.994 J/mol-K – p^o = 0.1 MPa



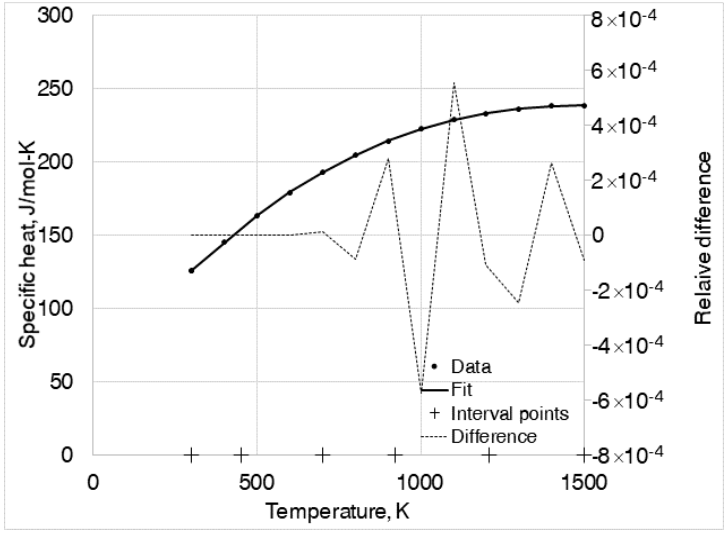


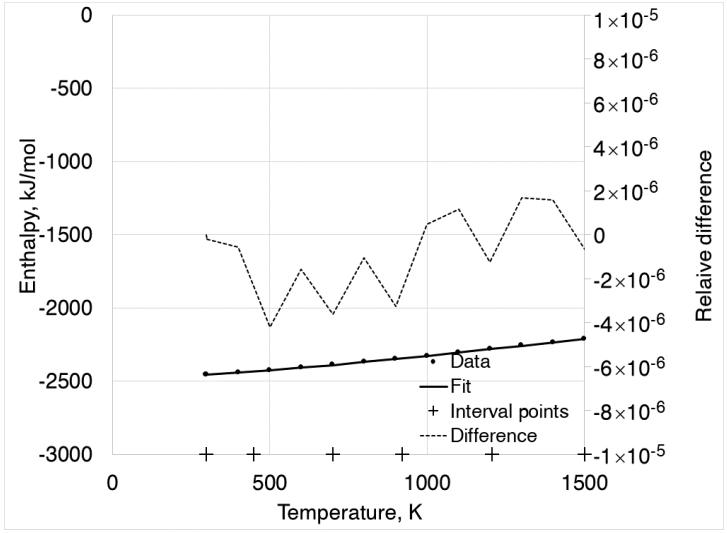


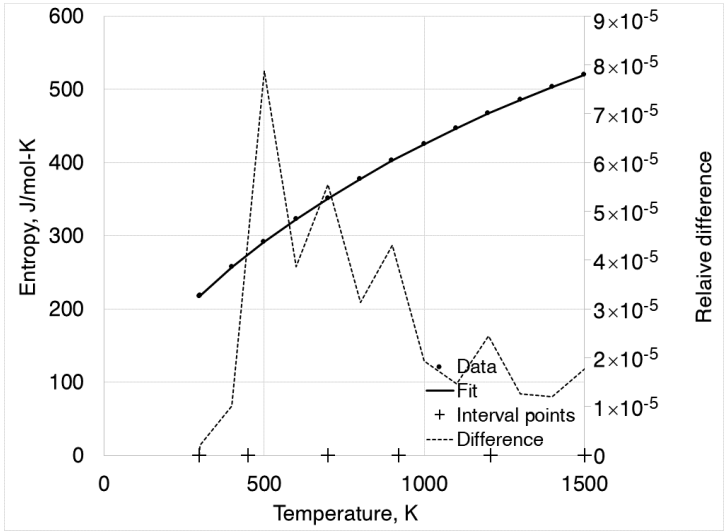


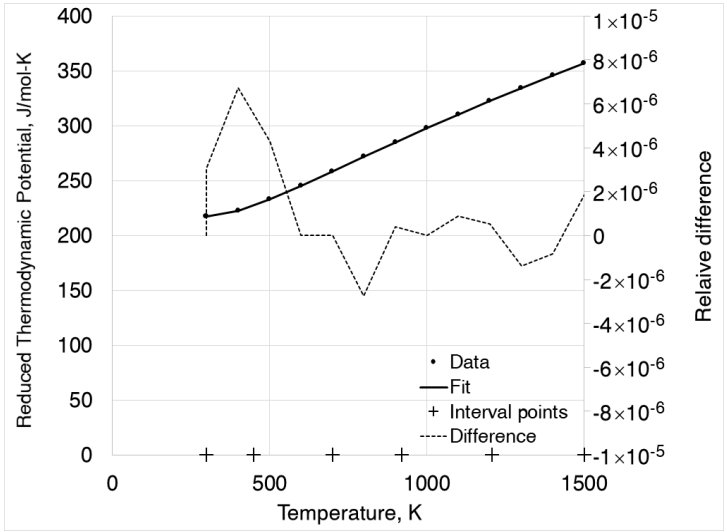
Trifluoroboroxin ($B_3O_3F_3$) – B3F3O3(cr)

MW = 137.42 g/mol – $\Delta_f H^o$(298.15 K) = -2453.916 kJ/mol – S^o(298.15 K) = 217.568 J/mol-K – p^o = 0.1 MPa



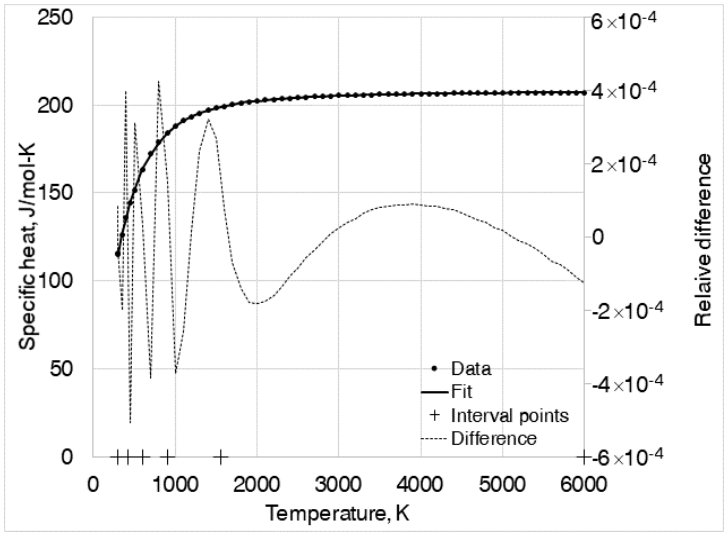


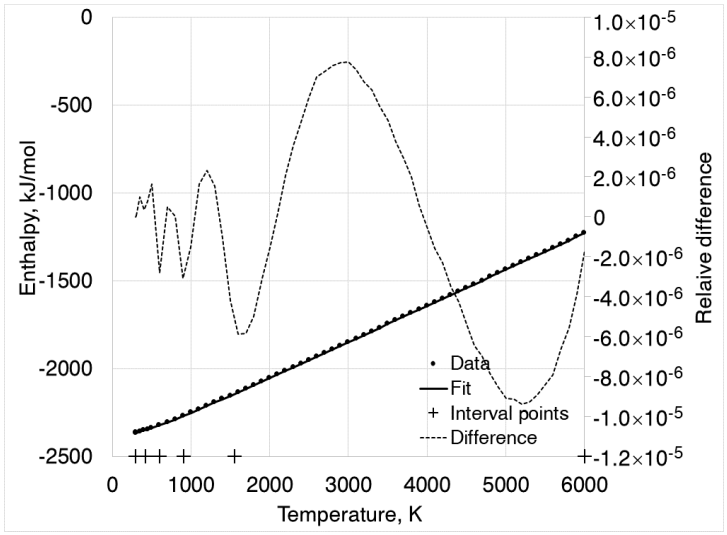


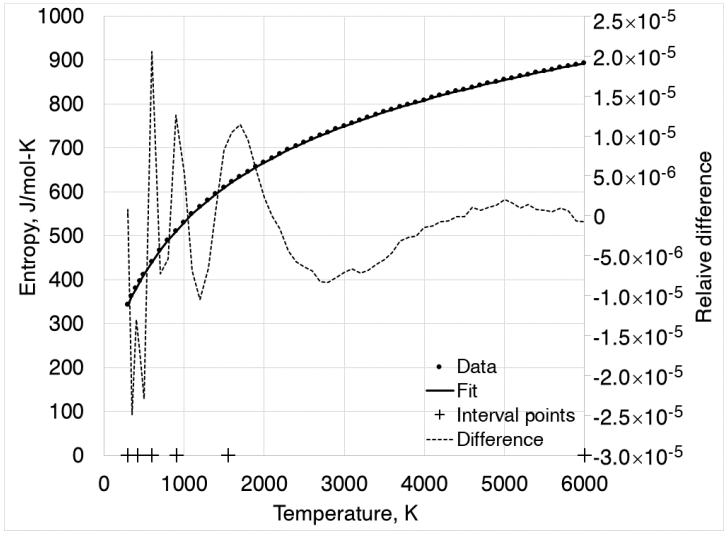


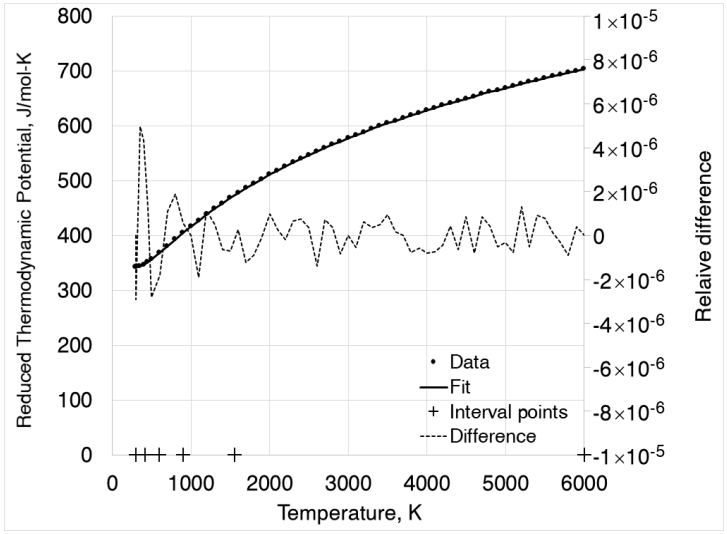
Trifluoroboroxin ($B_3O_3F_3$) – B3F3O3(g)

MW = 137.42 g/mol – $\Delta_f H^o$(298.15 K) = -2365.215 kJ/mol – S^o(298.15 K) = 342.454 J/mol-K – p^o = 0.1 MPa



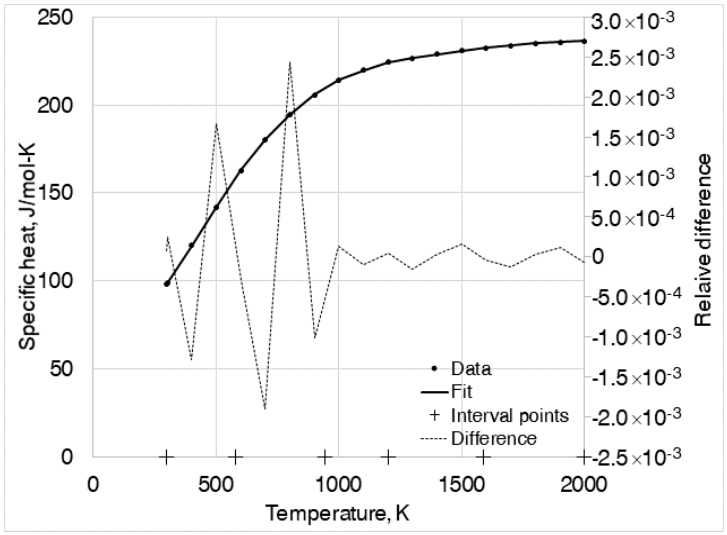


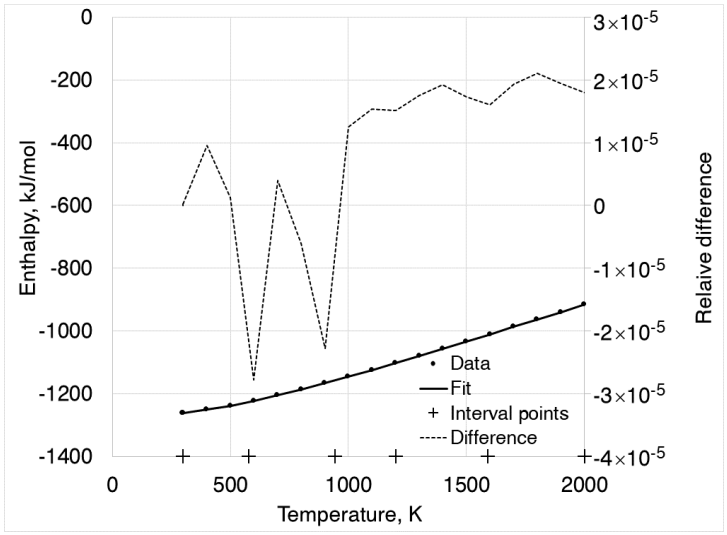


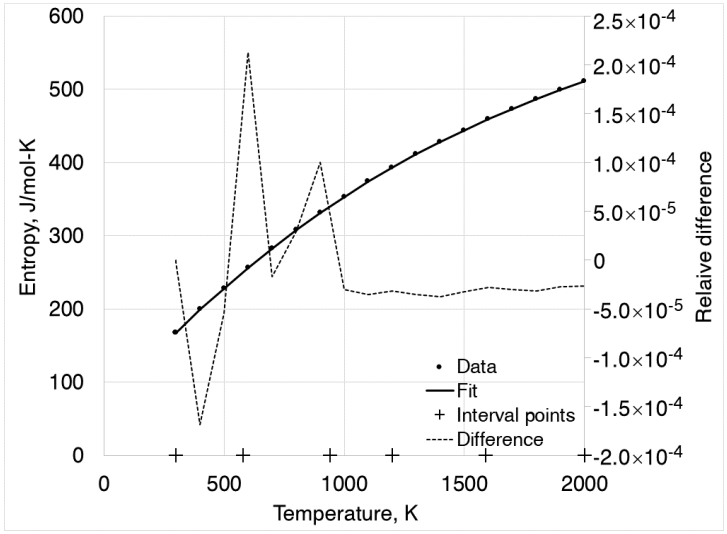


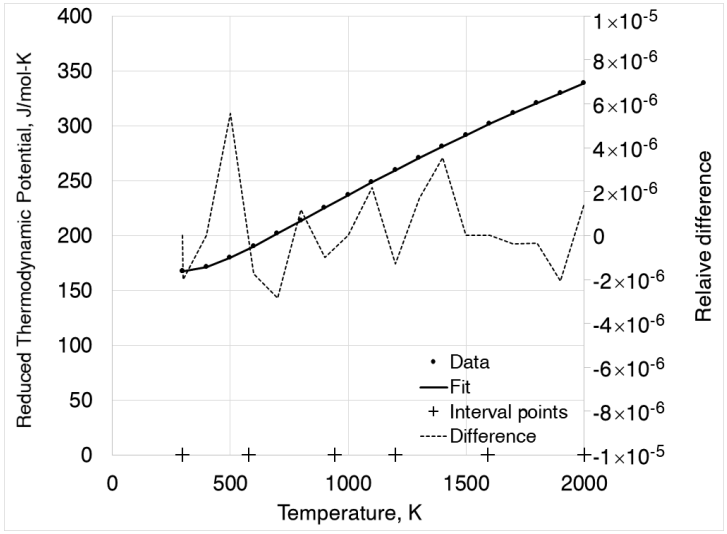
Boroxin ($B_3H_3O_3$) – B3H3O3(cr)

MW = 83.45 g/mol – $\Delta_f H^o$(298.15 K) = -1262.313 kJ/mol – S^o(298.15 K) = 167.36 J/mol-K – p^o = 0.1 MPa



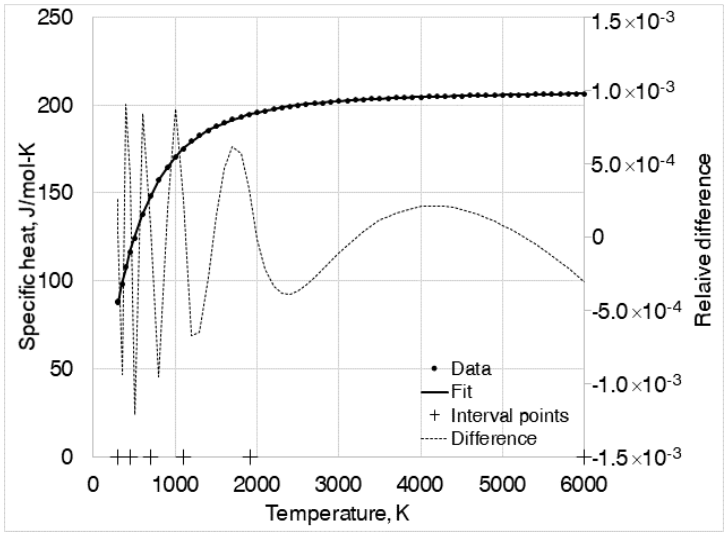


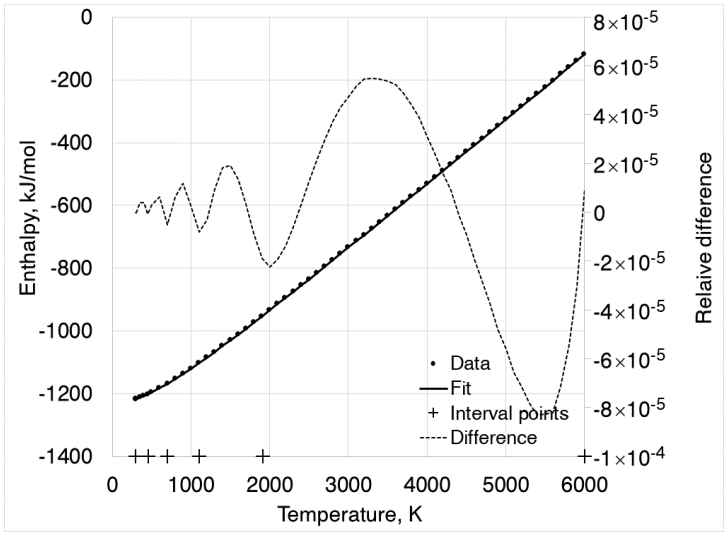


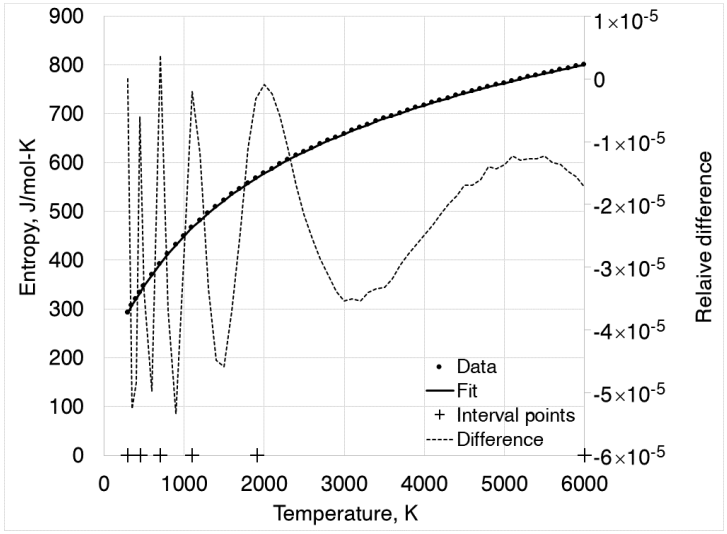


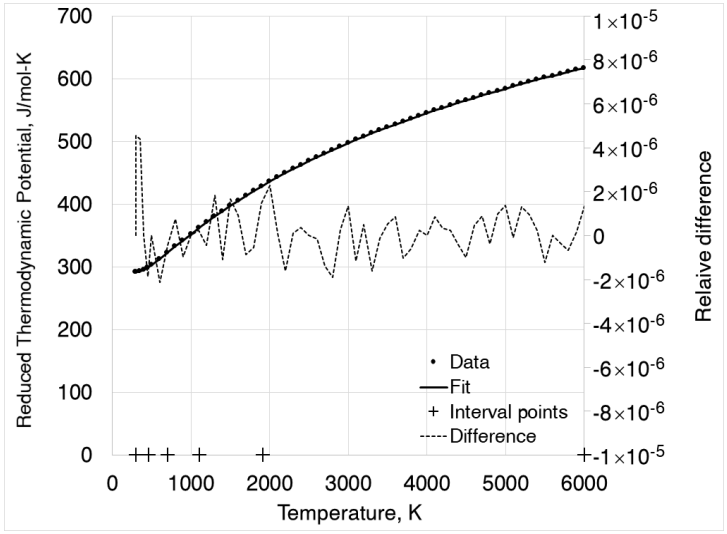
Boroxin ($B_3H_3O_3$) – B3H3O3(g)

MW = 83.45 g/mol – $\Delta_f H^o$(298.15 K) = -1217.544 kJ/mol – S^o(298.15 K) = 291.907 J/mol-K – p^o = 0.1 MPa



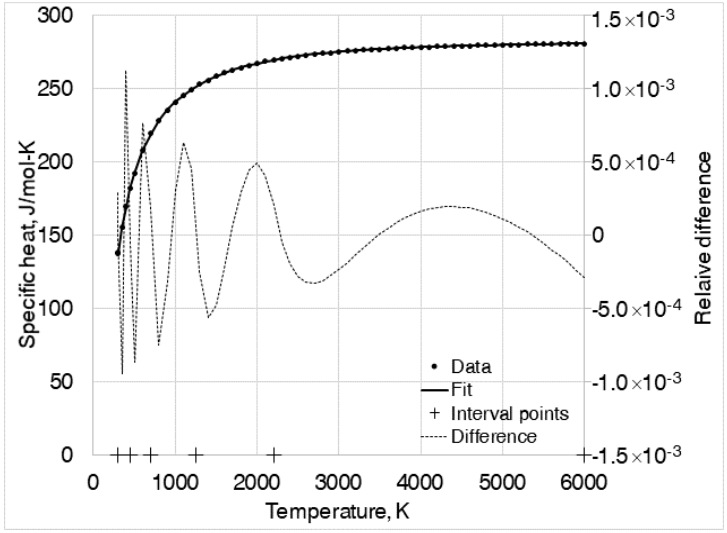


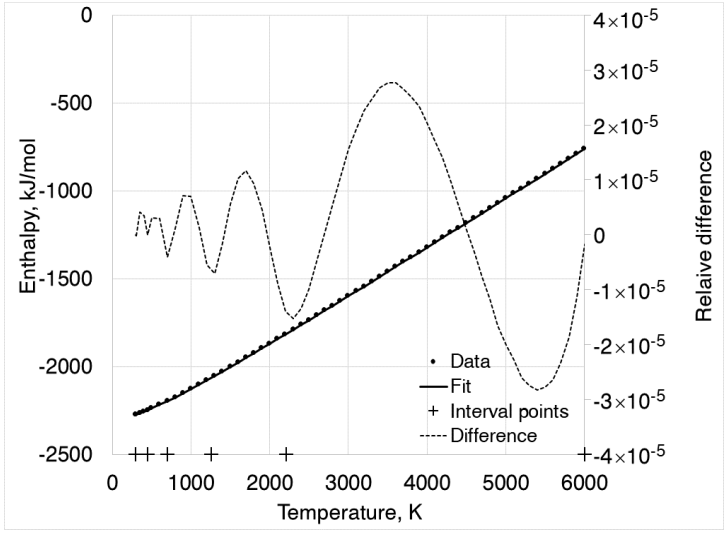


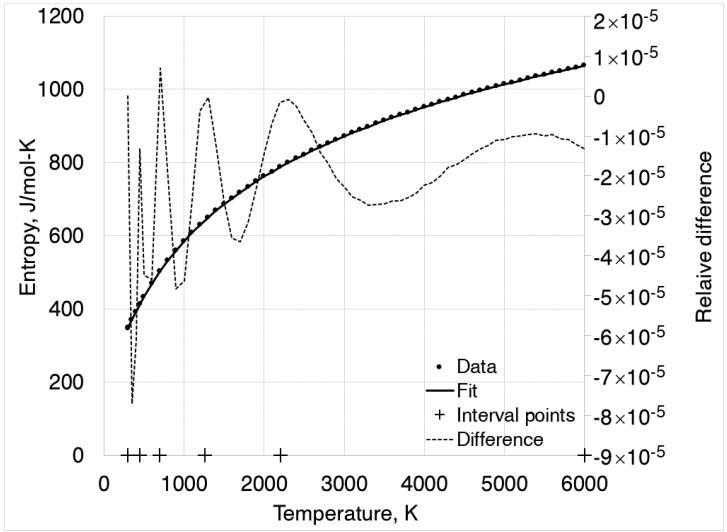


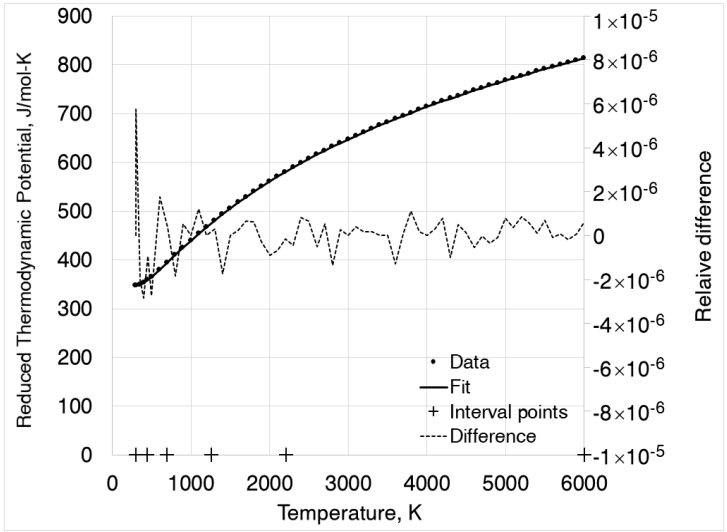
Boric Acid ($(HBO_2)_3$) – B3H3O6(g)

MW = 131.45 g/mol – $\Delta_f H^o$(298.15 K) = -2271.912 kJ/mol – S^o(298.15 K) = 347.567 J/mol-K – p^o = 0.1 MPa



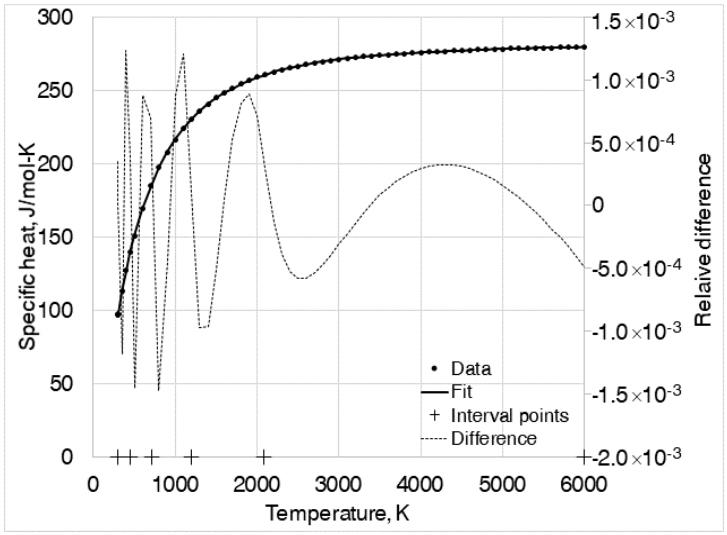


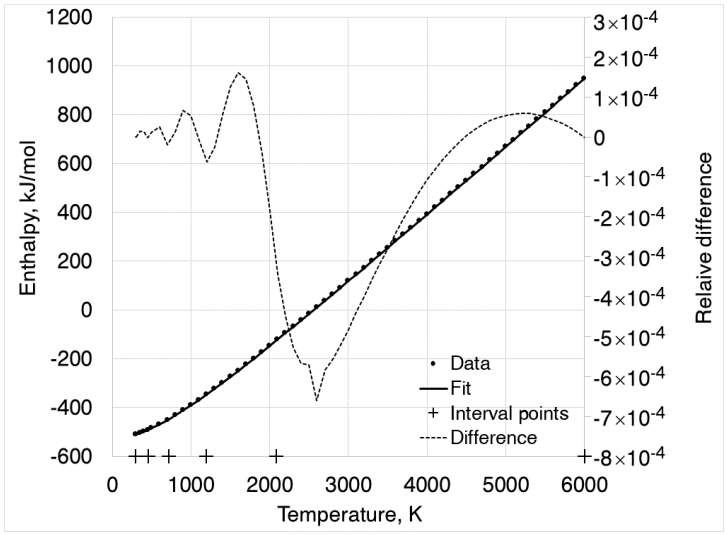


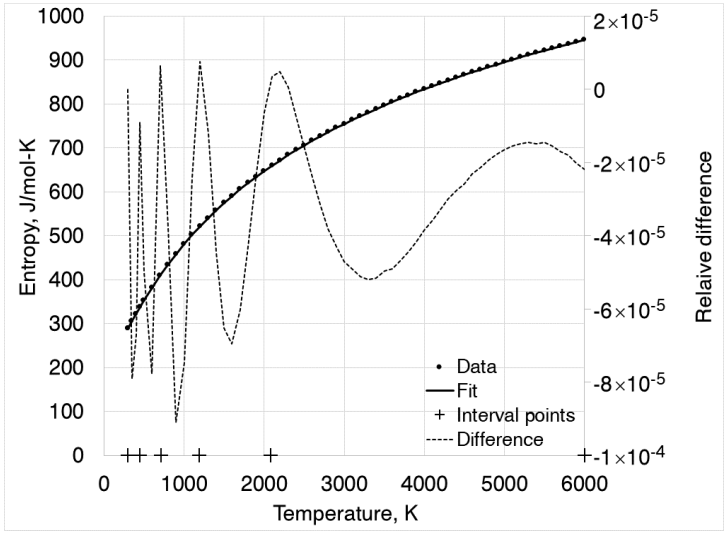


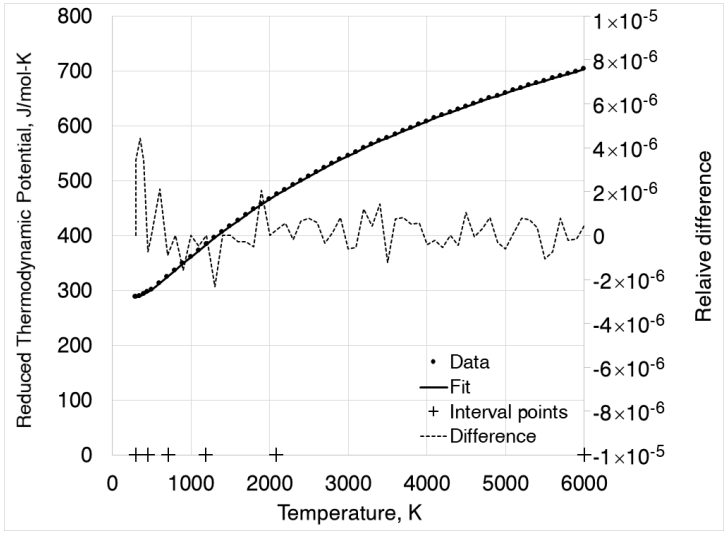
Borazine ($B_3H_6N_3$) – B3H6N3(g)

MW = 80.5 g/mol – $\Delta_f H^o$(298.15 K) = -510.03 kJ/mol – S^o(298.15 K) = 288.723 J/mol-K – p^o = 0.1 MPa



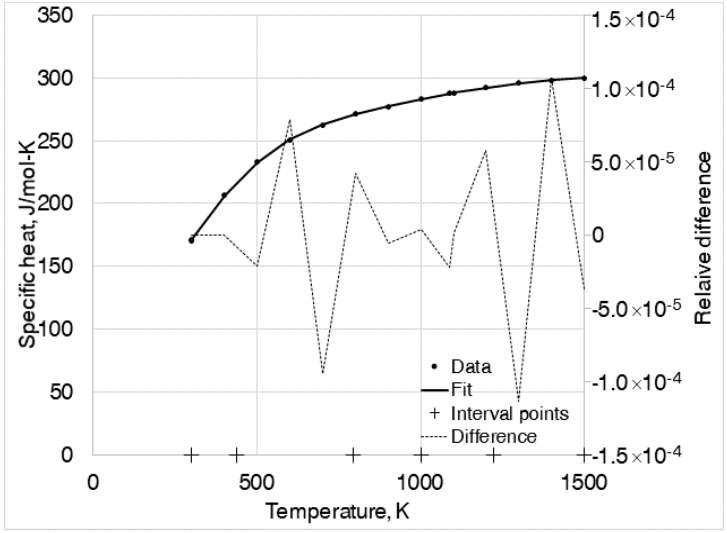


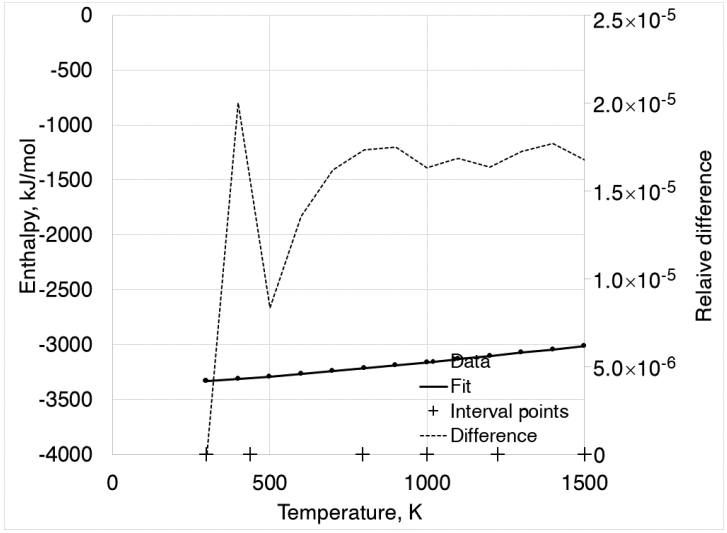


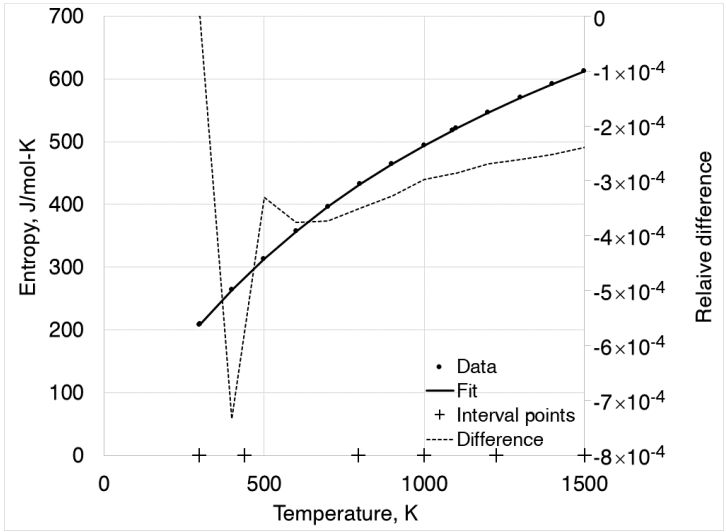


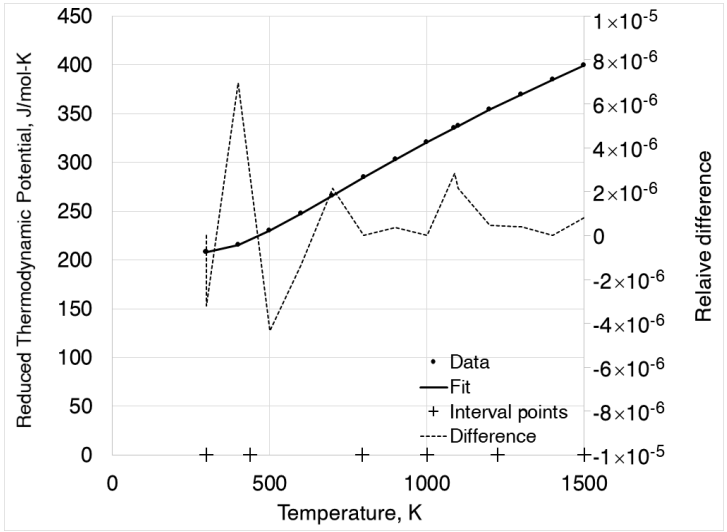
Potassium Borate ($K_2B_4O_7$) – B4K2O7(cr)

MW = 233.43 g/mol – $\Delta_f H^o$(298.15 K) = -3334.23 kJ/mol – S^o(298.15 K) = 208.363 J/mol-K – p^o = 0.1 MPa



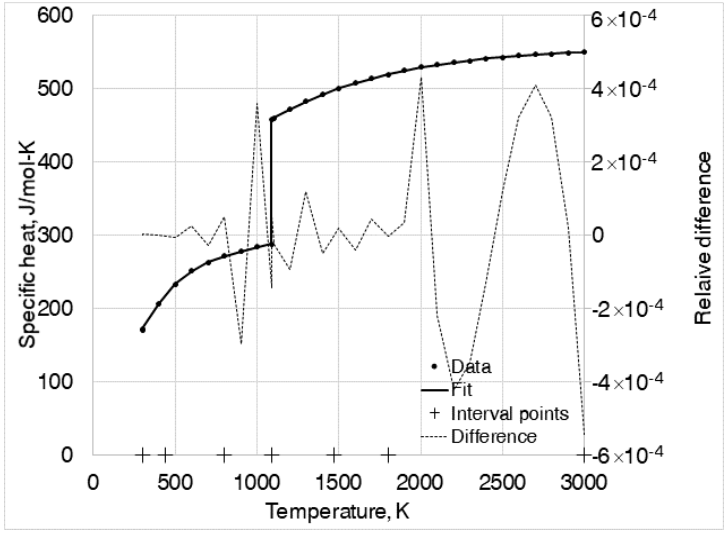


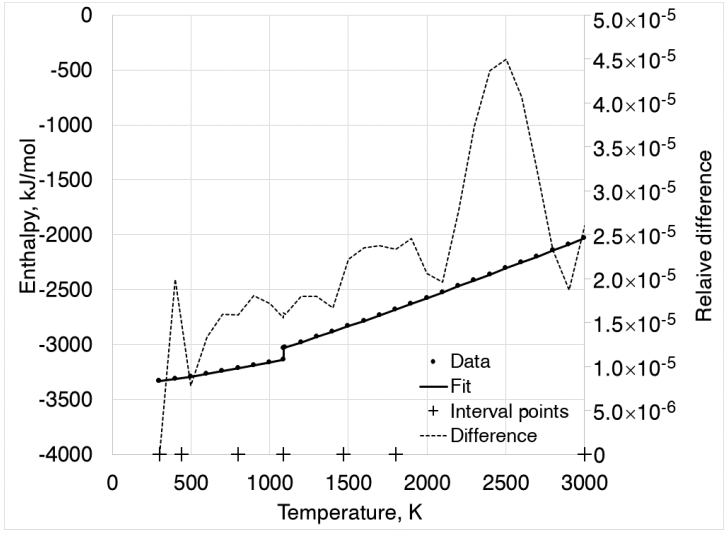


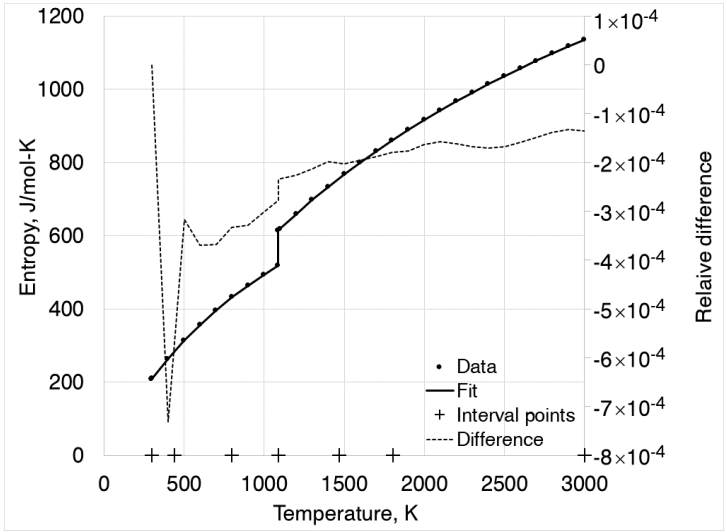


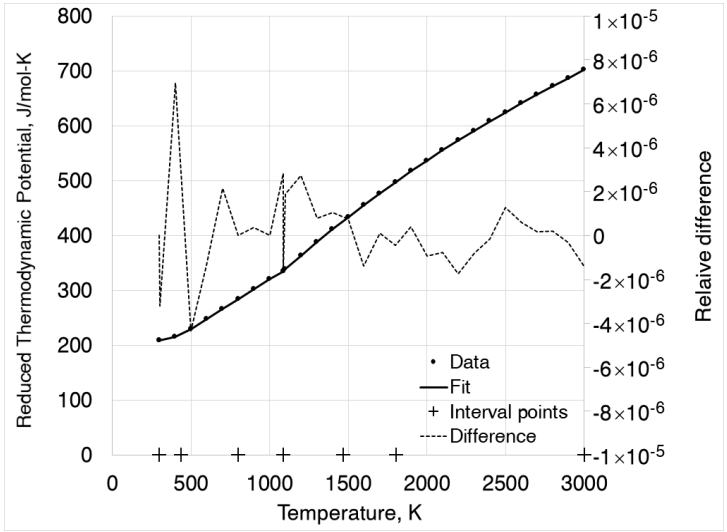
Potassium Borate ($K_2B_4O_7$) – B4K2O7(cr,l)

MW = 233.43 g/mol – $\Delta_f H^o$(298.15 K) = -3334.23 kJ/mol – S^o(298.15 K) = 208.363 J/mol-K – p^o = 0.1 MPa



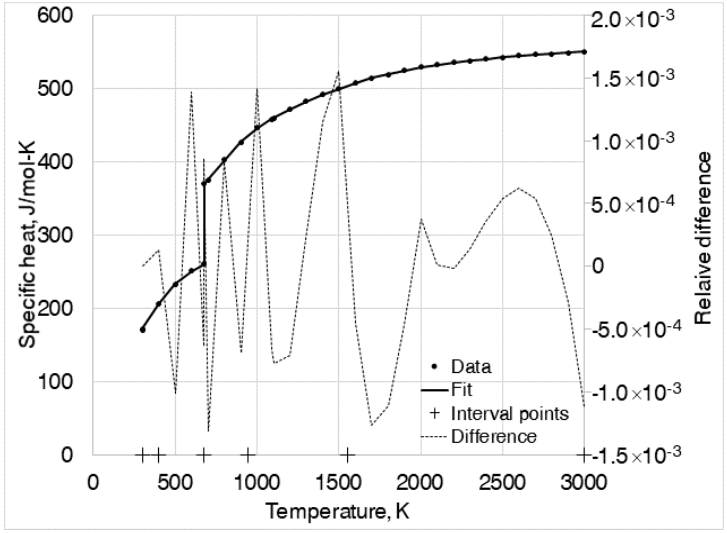


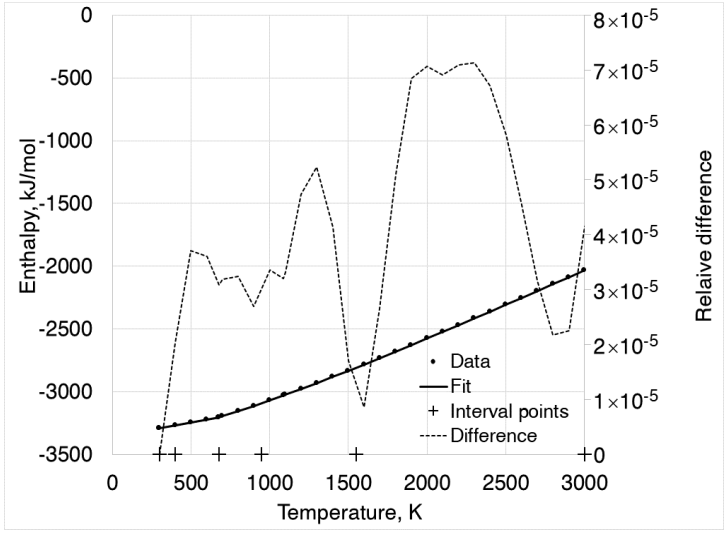


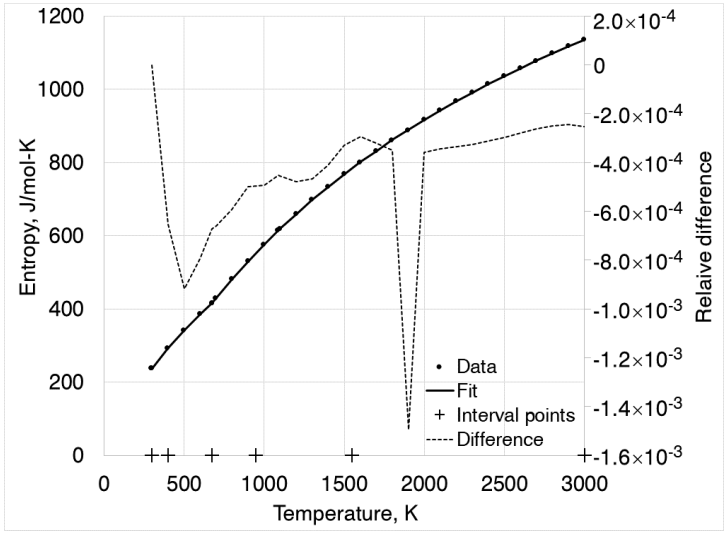


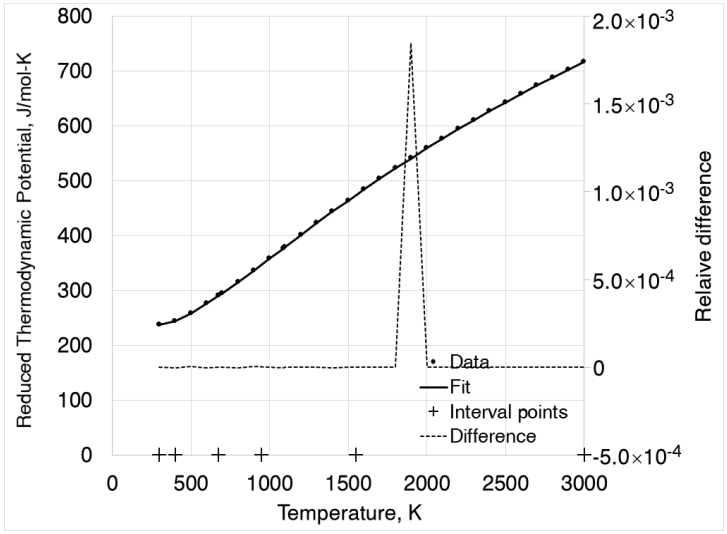
Potassium Borate ($K_2B_4O_7$) – B4K2O7(l)

MW = 233.43 g/mol – $\Delta_f H^o$(298.15 K) = -3289.042 kJ/mol – S^o(298.15 K) = 237.12 J/mol-K – p^o = 0.1 MPa



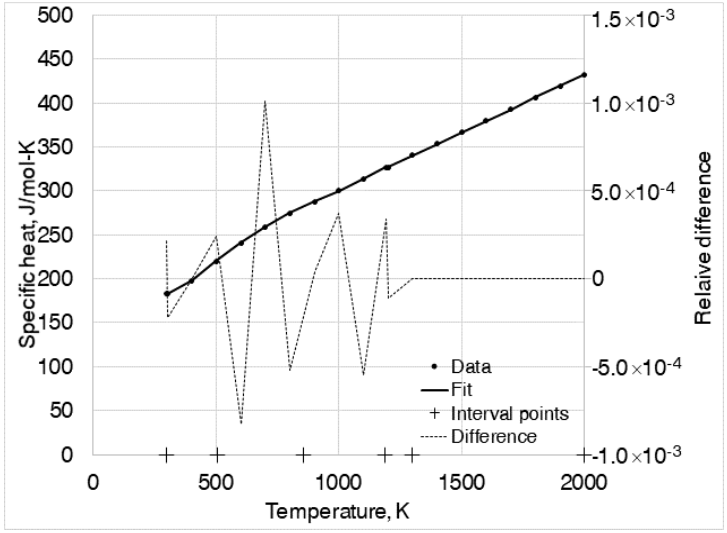


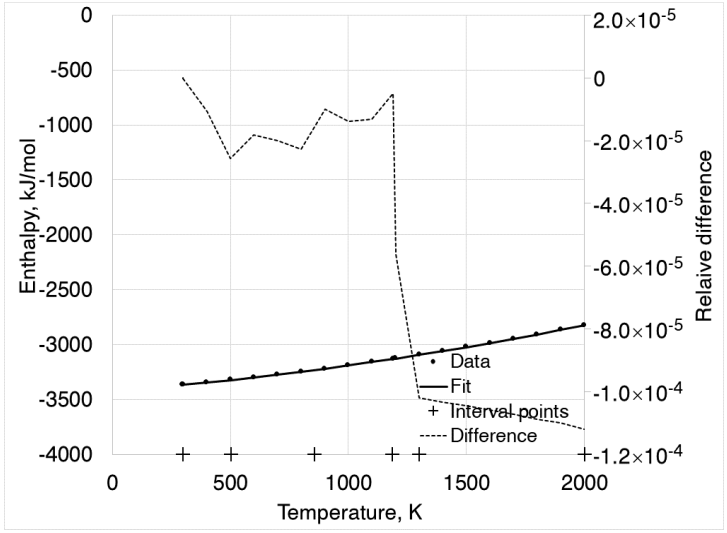


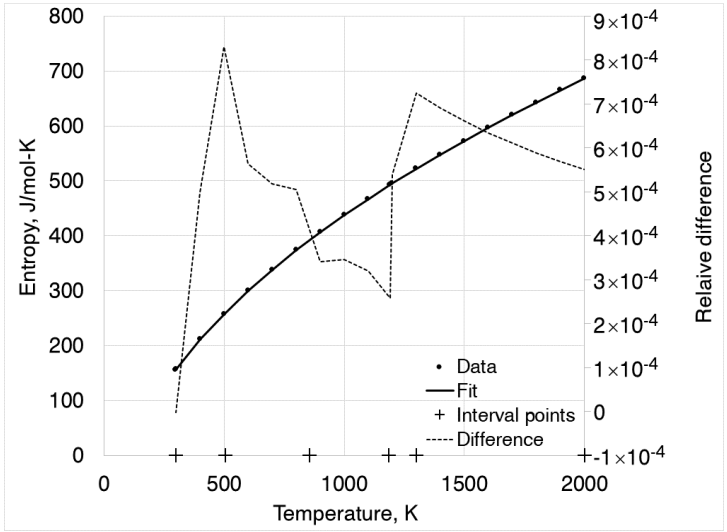


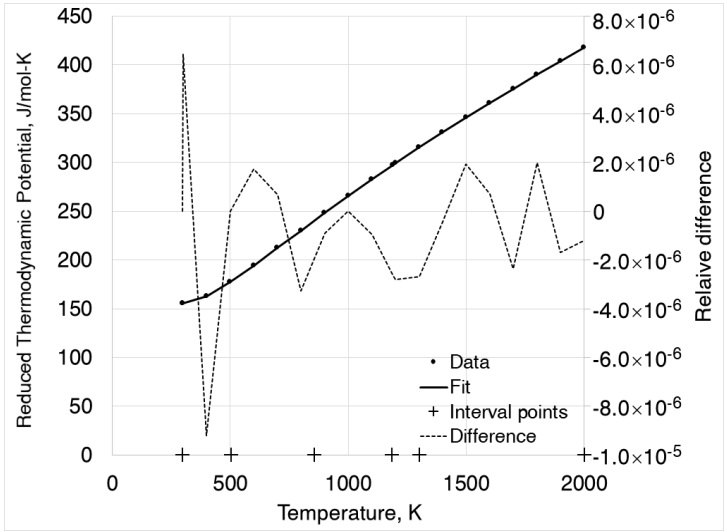
Lithium Borate ($Li_2B_4O_7$) – B4Li2O7(cr)

MW = 169.12 g/mol – $\Delta_f H^o$(298.15 K) = -3362.262 kJ/mol – S^o(298.15 K) = 155.645 J/mol-K – p^o = 0.1 MPa



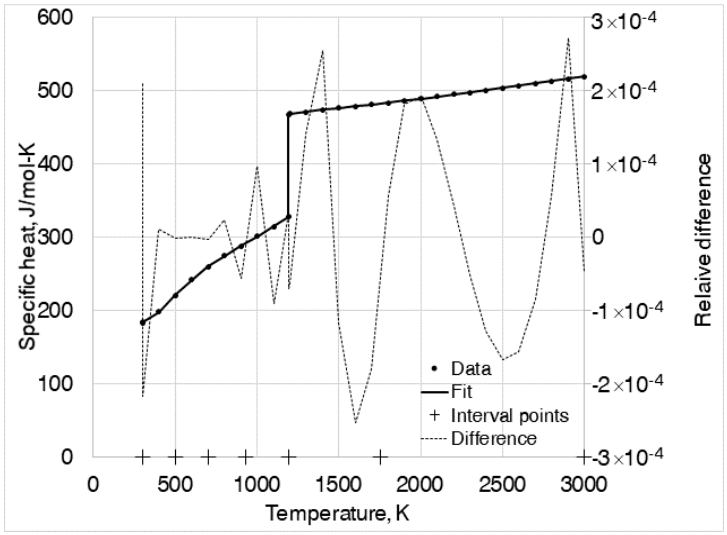


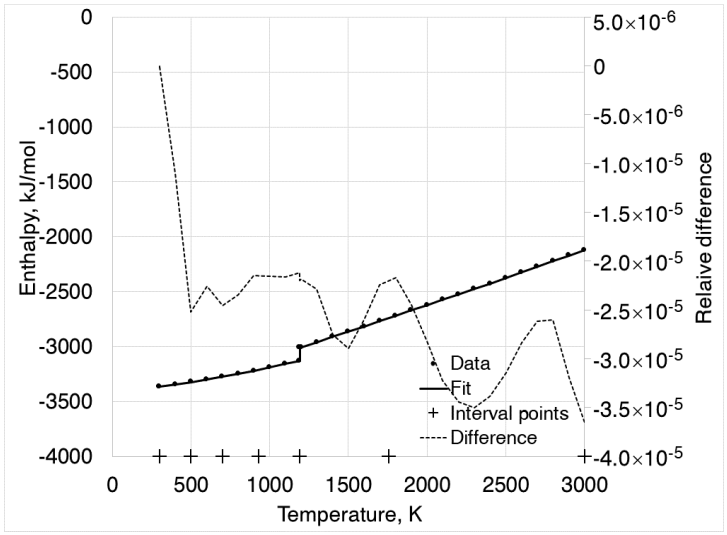


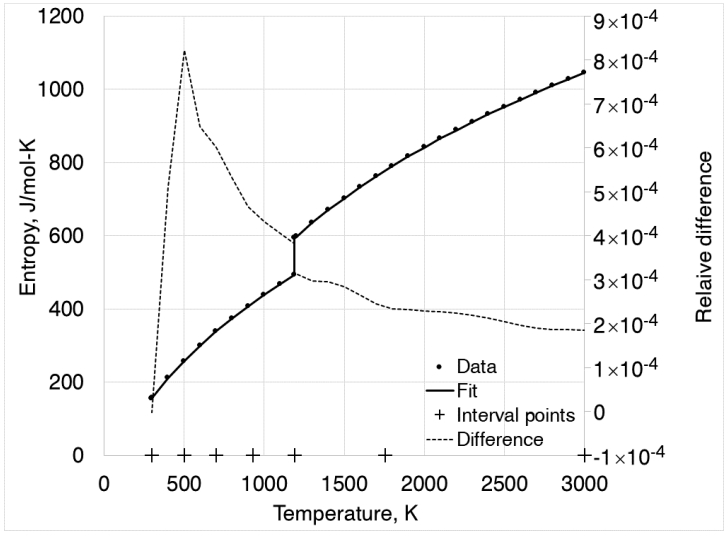


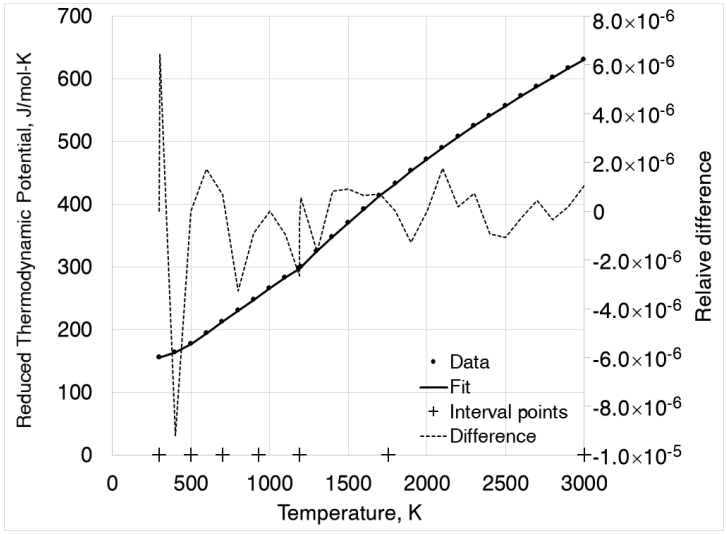
Lithium Borate ($Li_2B_4O_7$) – B4Li2O7(cr,l)

MW = 169.12 g/mol – $\Delta_f H^o$(298.15 K) = -3362.262 kJ/mol – S^o(298.15 K) = 155.645 J/mol-K – p^o = 0.1 MPa



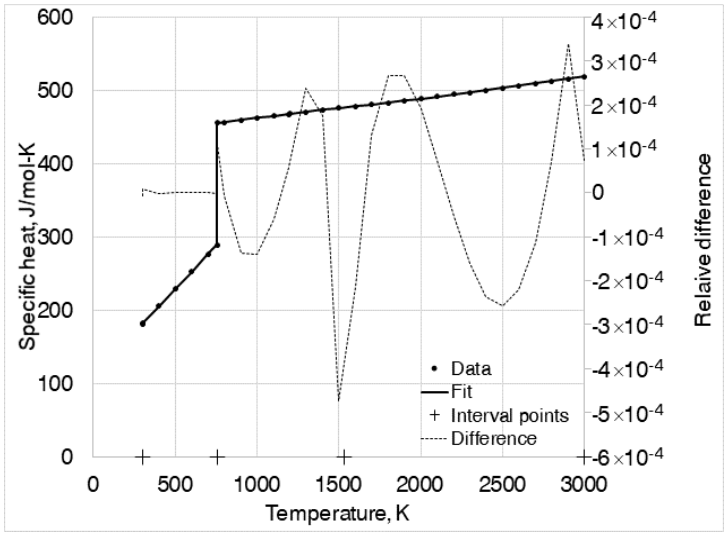


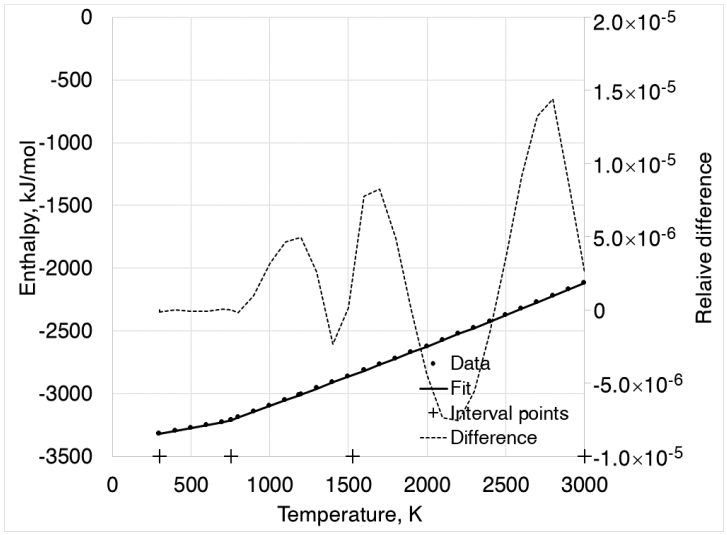


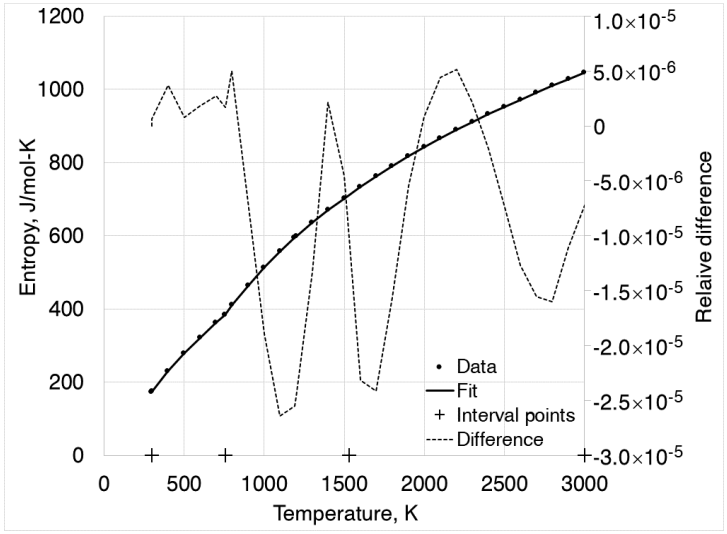


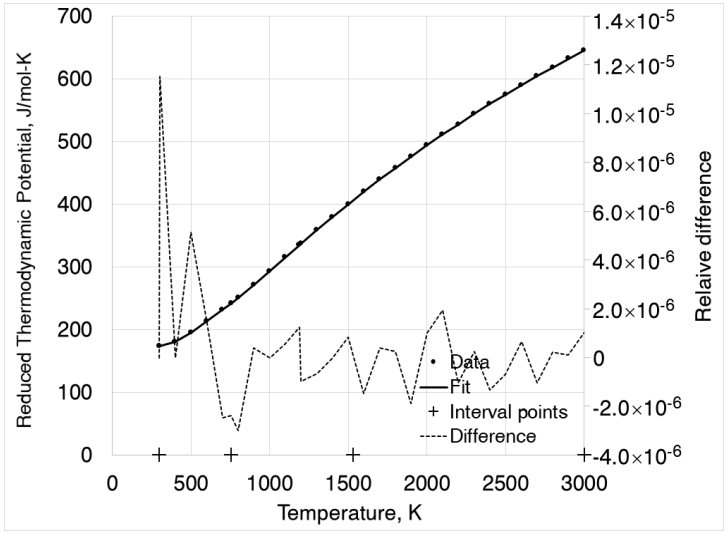
Lithium Borate ($Li_2B_4O_7$) – B4Li2O7(l)

MW = 169.12 g/mol – $\Delta_f H^o$(298.15 K) = -3317.715 kJ/mol – S^o(298.15 K) = 173.375 J/mol-K – p^o = 0.1 MPa



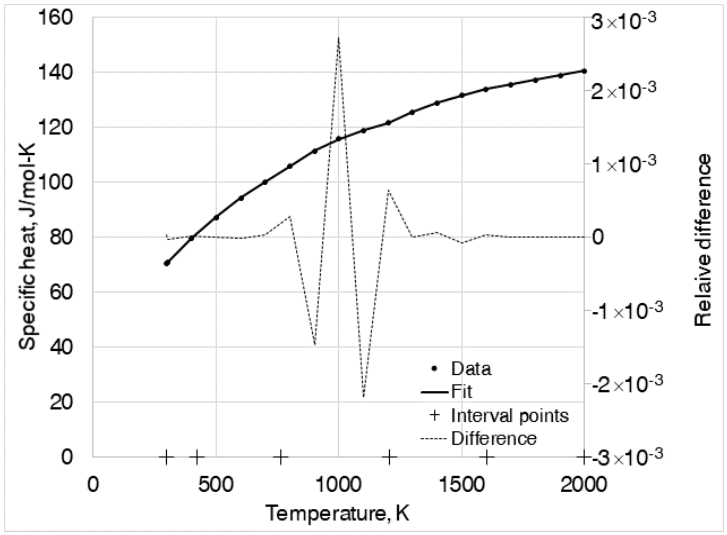


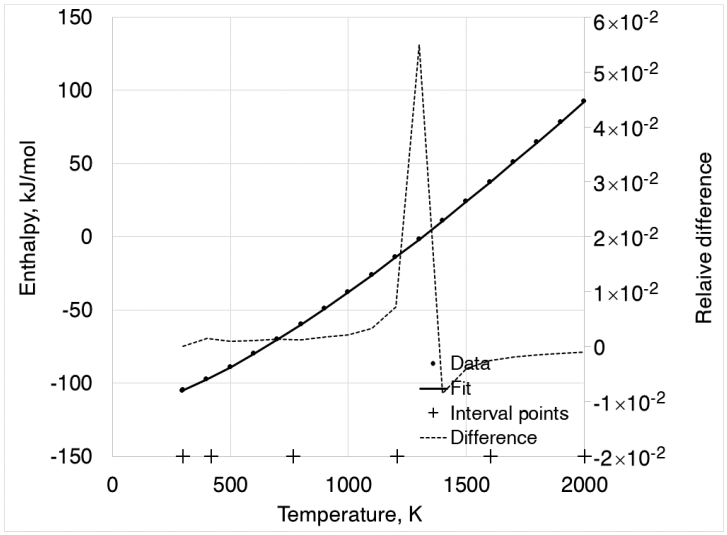


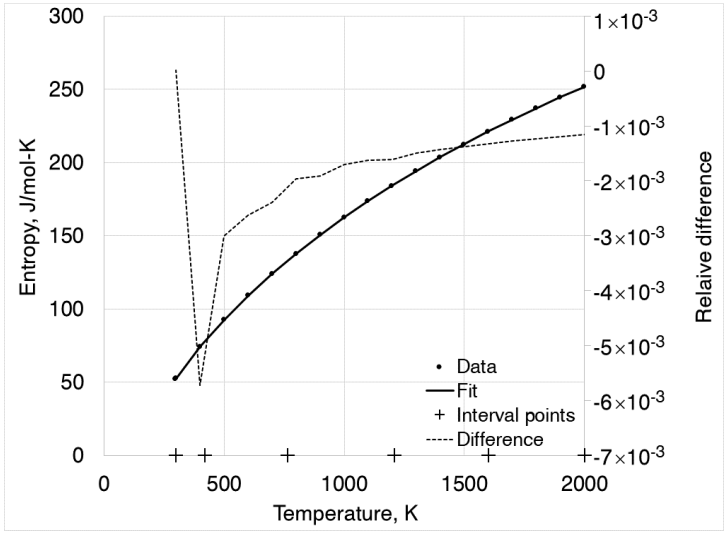


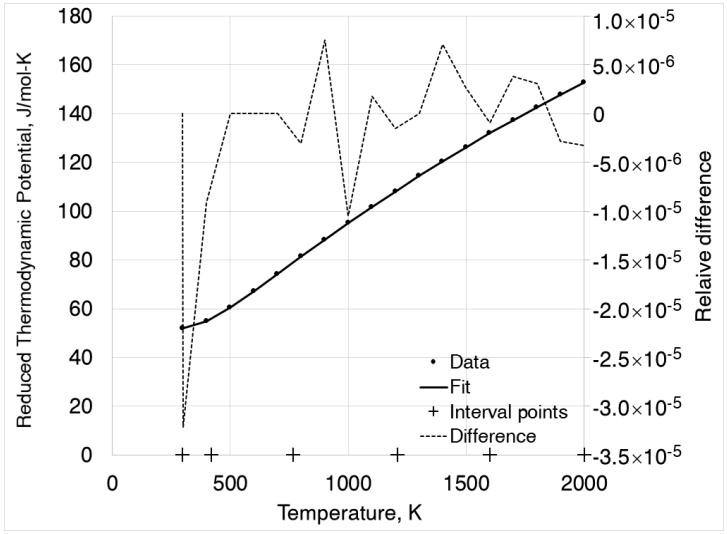
Magnesium Boride (MgB_4) – B4Mg1(cr)

MW = 67.54 g/mol – $\Delta_f H^o$(298.15 K) = -105.018 kJ/mol – S^o(298.15 K) = 51.923 J/mol-K – p^o = 0.1 MPa



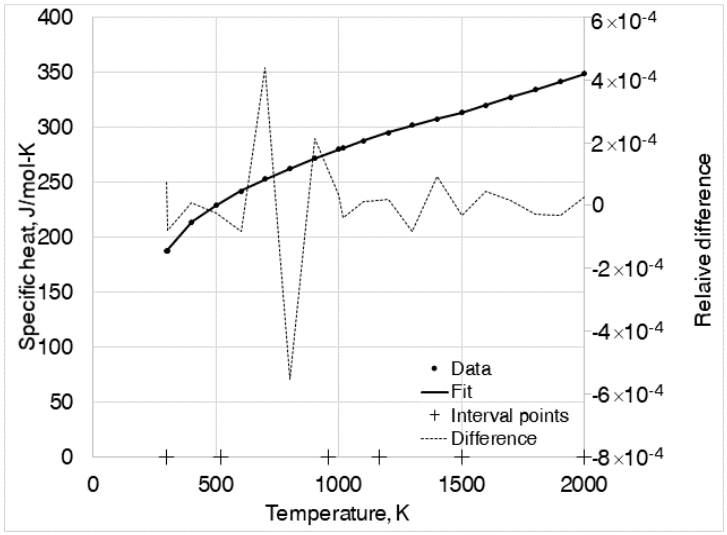


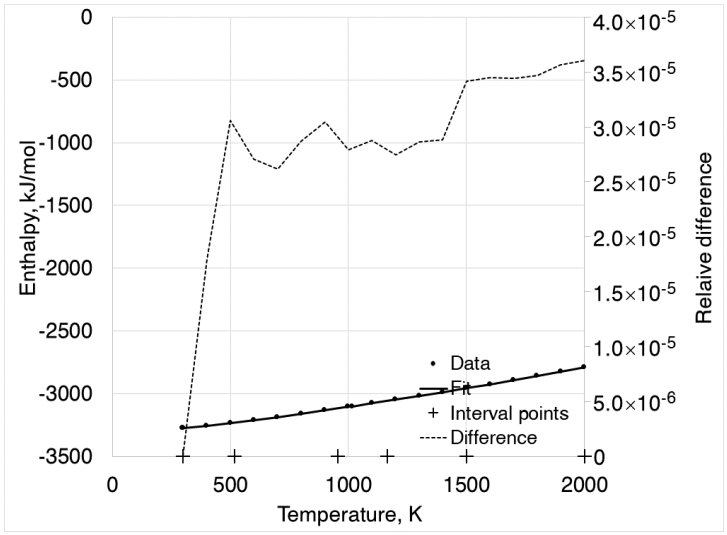


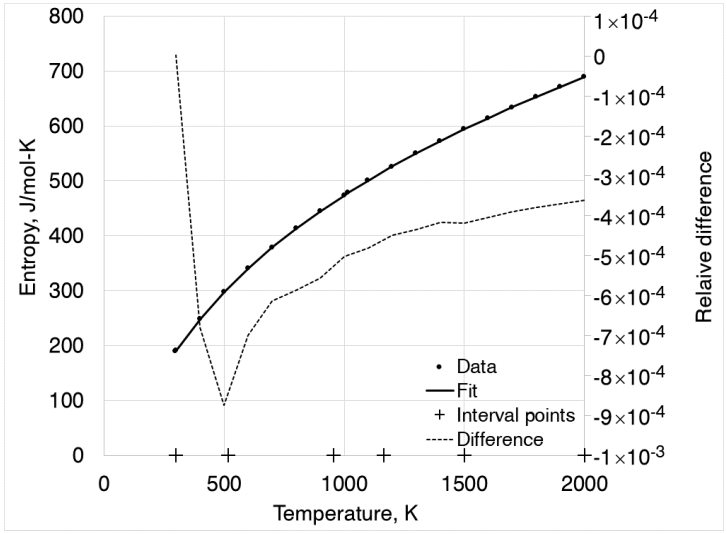


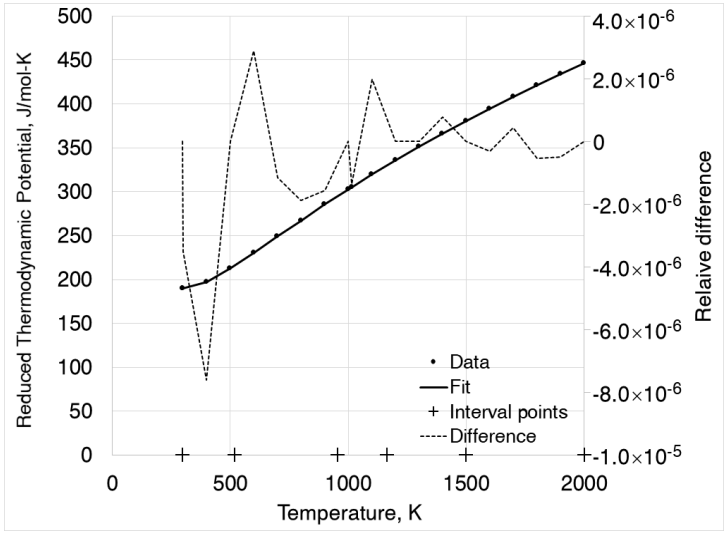
Sodium Borate ($Na_2B_4O_7$) – B4Na2O7(cr)

MW = 201.22 g/mol – $\Delta_f H^o$(298.15 K) = -3276.741 kJ/mol – S^o(298.15 K) = 189.49 J/mol-K – p^o = 0.1 MPa



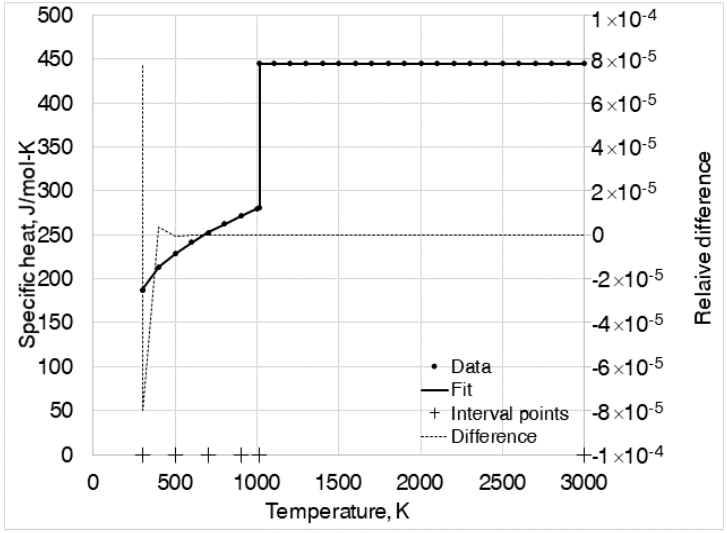


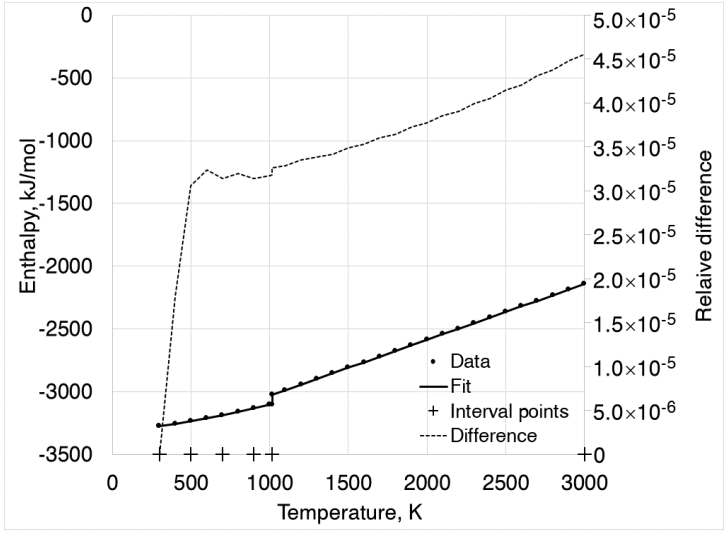


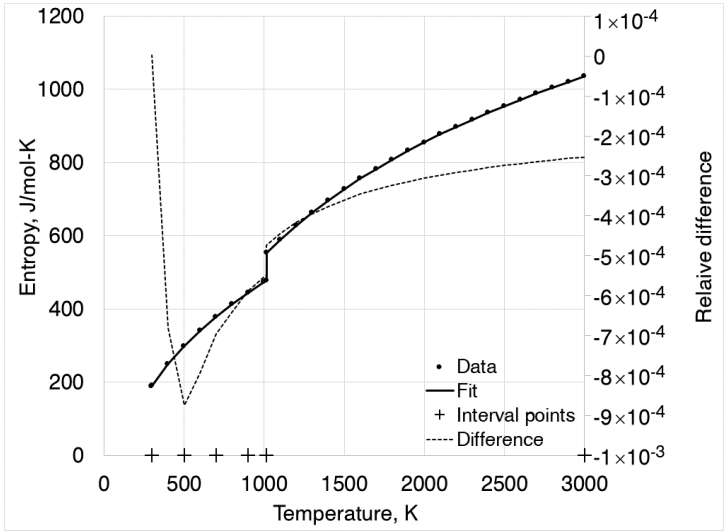


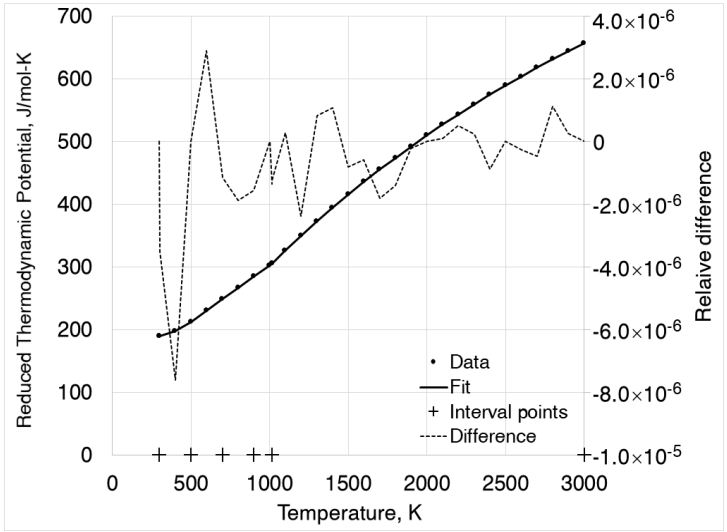
Sodium Borate ($Na_2B_4O_7$) – B4Na2O7(cr,l)

MW = 201.22 g/mol – $\Delta_f H^o$(298.15 K) = -3276.741 kJ/mol – S^o(298.15 K) = 189.49 J/mol-K – p^o = 0.1 MPa



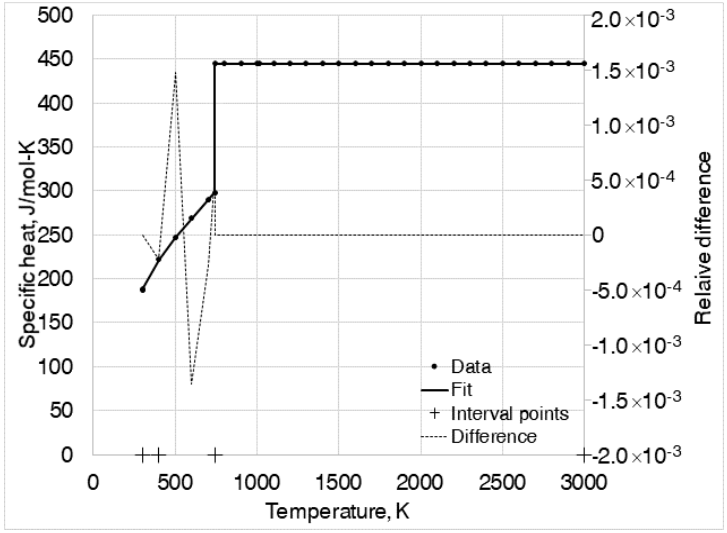


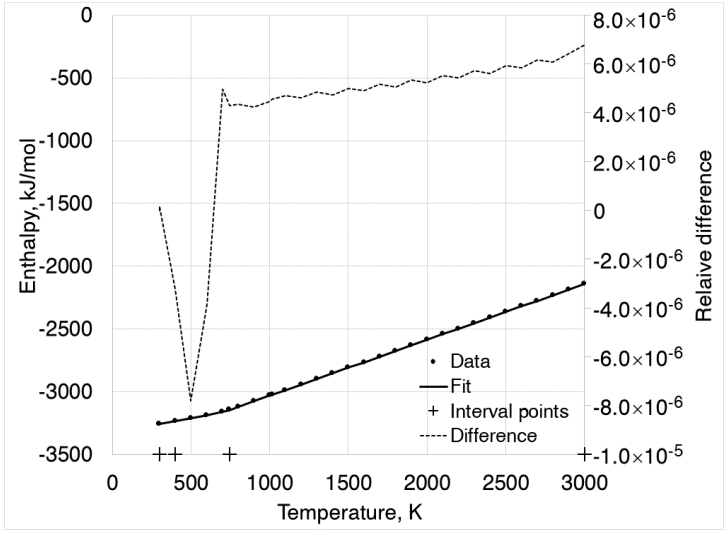


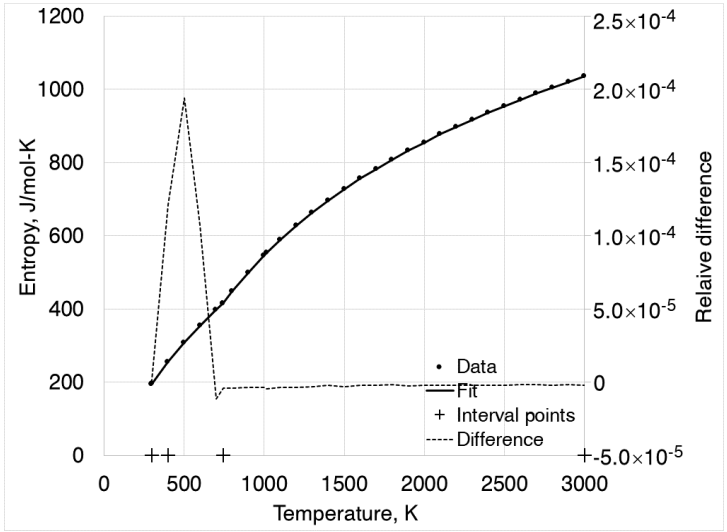


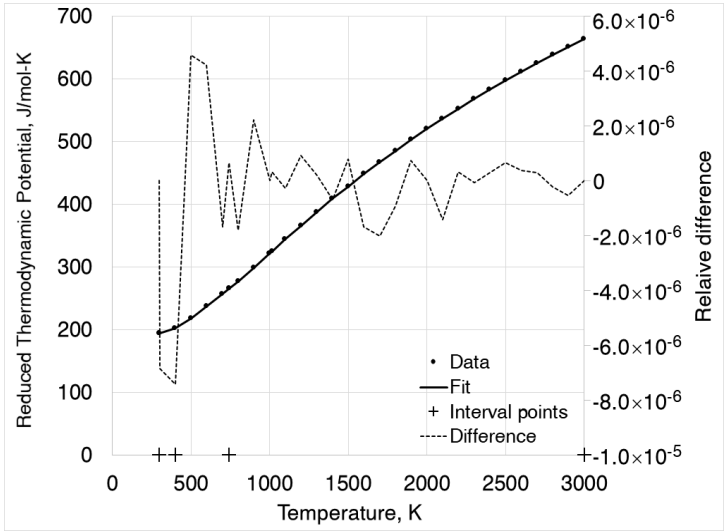
Sodium Borate ($Na_2B_4O_7$) – B4Na2O7(l)

MW = 201.22 g/mol – $\Delta_f H^o$(298.15 K) = -3256.658 kJ/mol – S^o(298.15 K) = 194.636 J/mol-K – p^o = 0.1 MPa



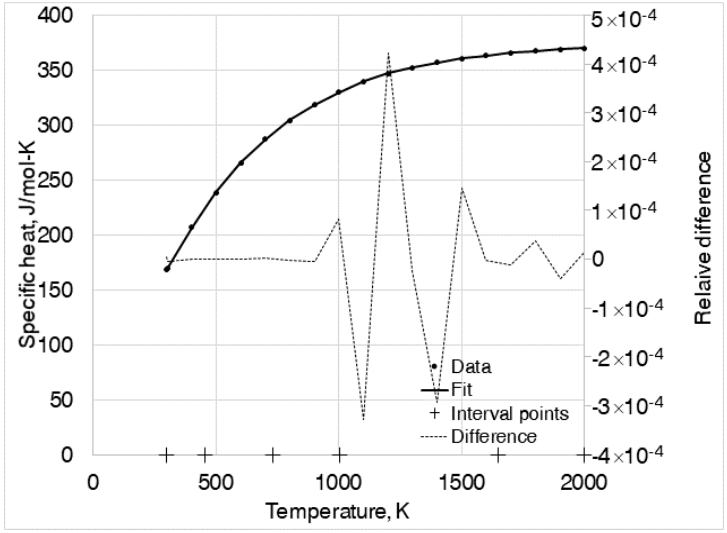


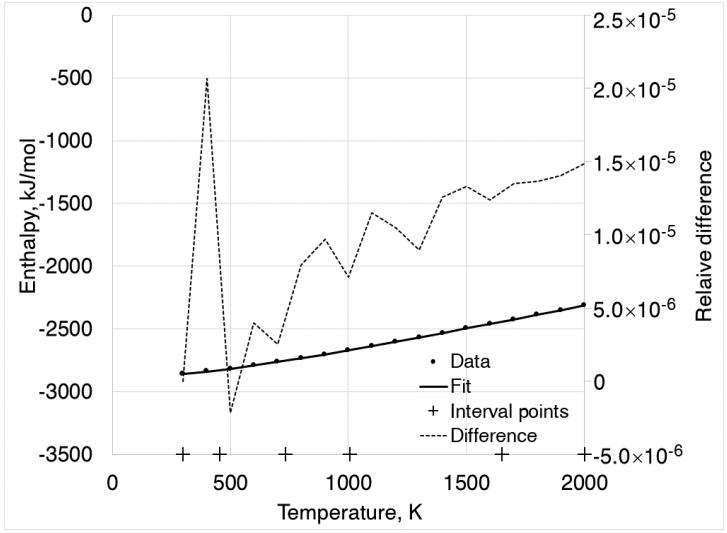


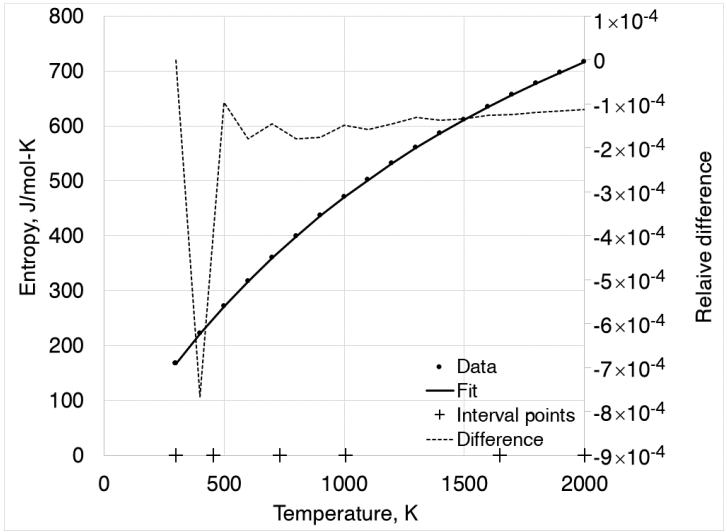


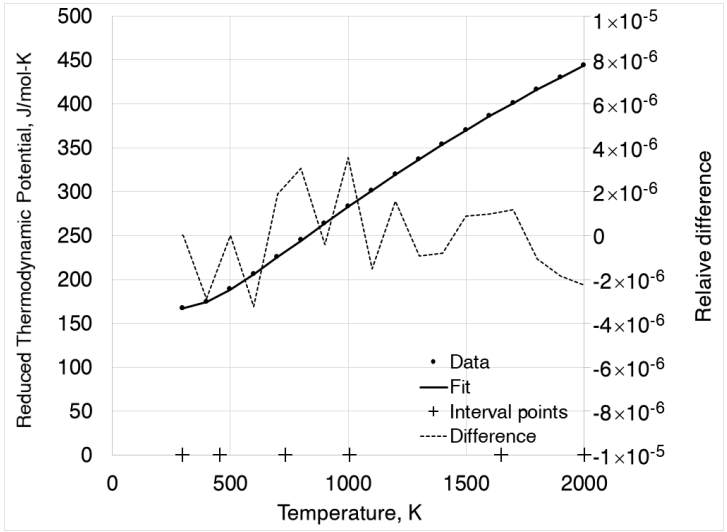
Lead Borate (PbB_4O_7) – B4O7Pb1(cr)

MW = 362.4 g/mol – $\Delta_f H^o$(298.15 K) = -2857.672 kJ/mol – S^o(298.15 K) = 166.942 J/mol-K – p^o = 0.1 MPa



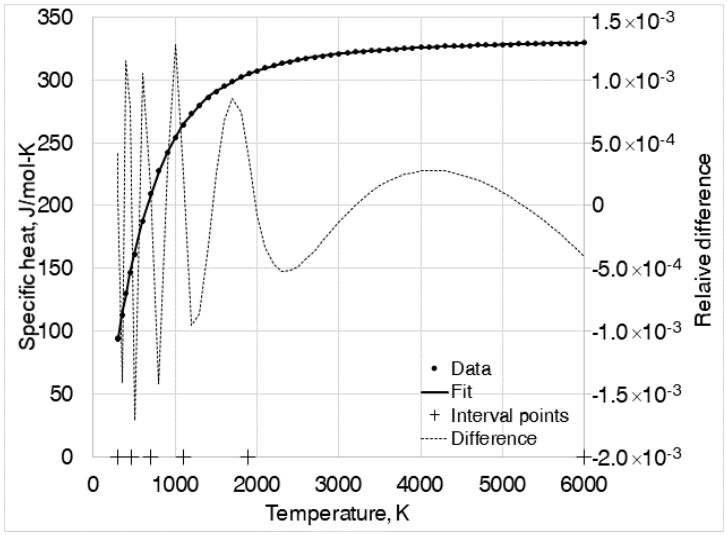


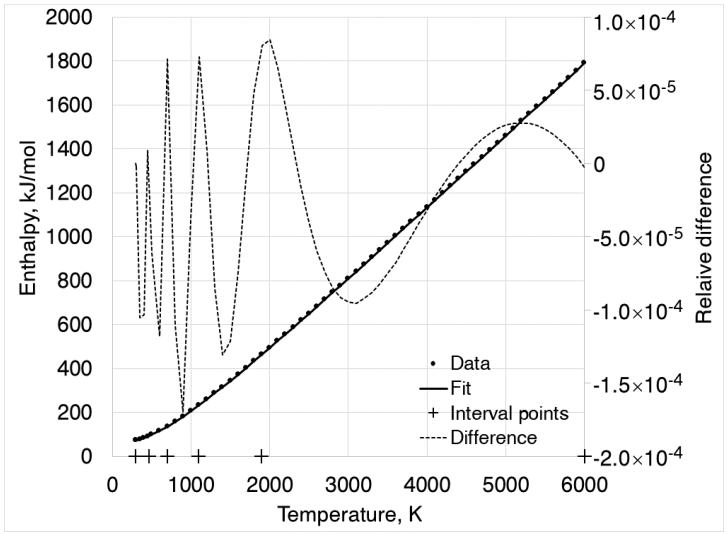


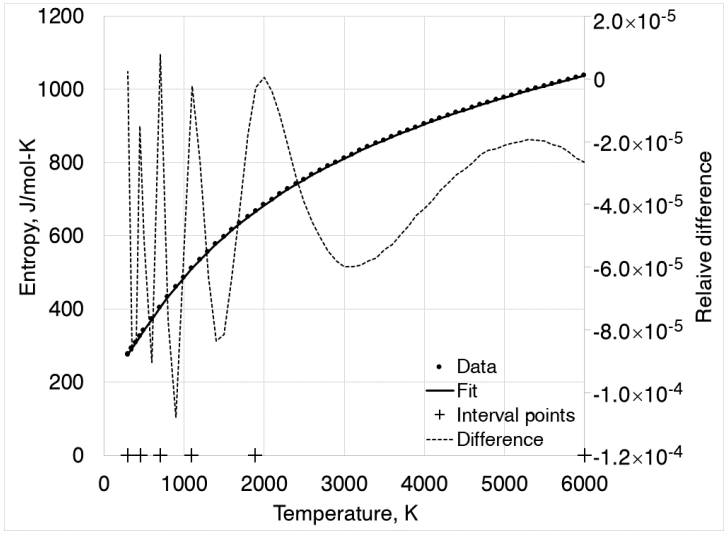


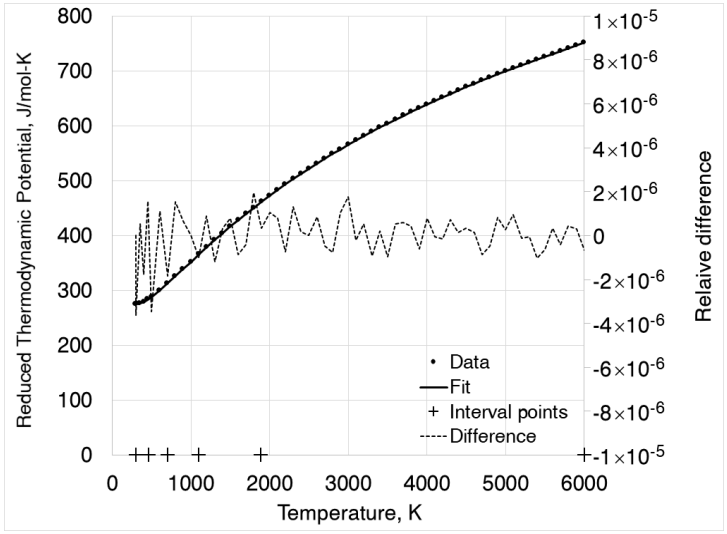
Pentaborane (B_5H_9) – B5H9(g)

MW = 63.12 g/mol – $\Delta_f H^o$(298.15 K) = 73.22 kJ/mol – S^o(298.15 K) = 275.41 J/mol-K – p^o = 0.1 MPa



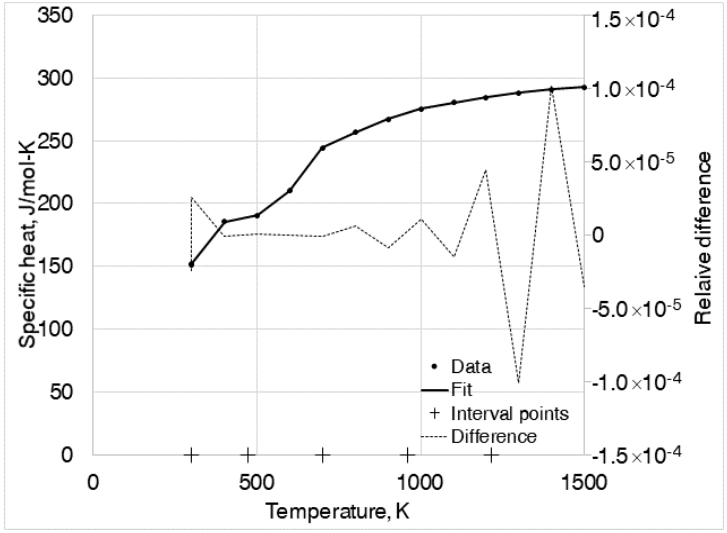


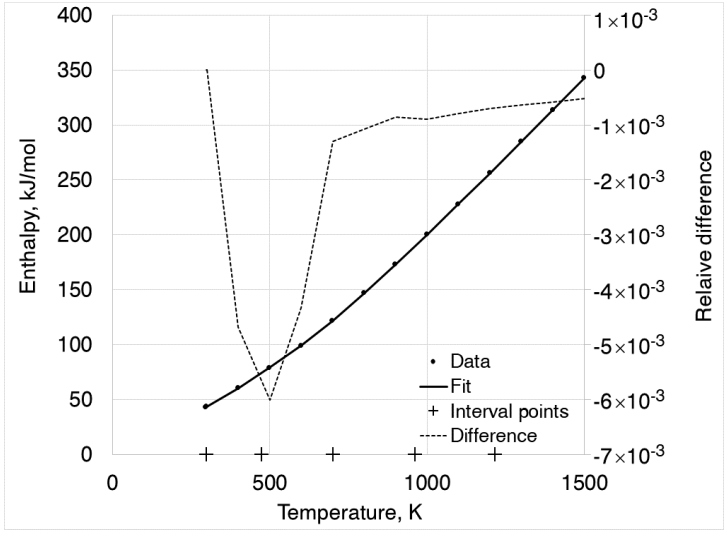


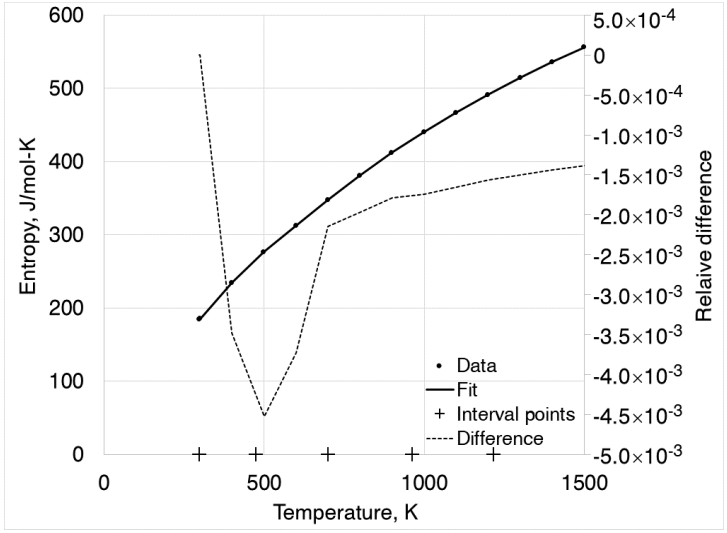


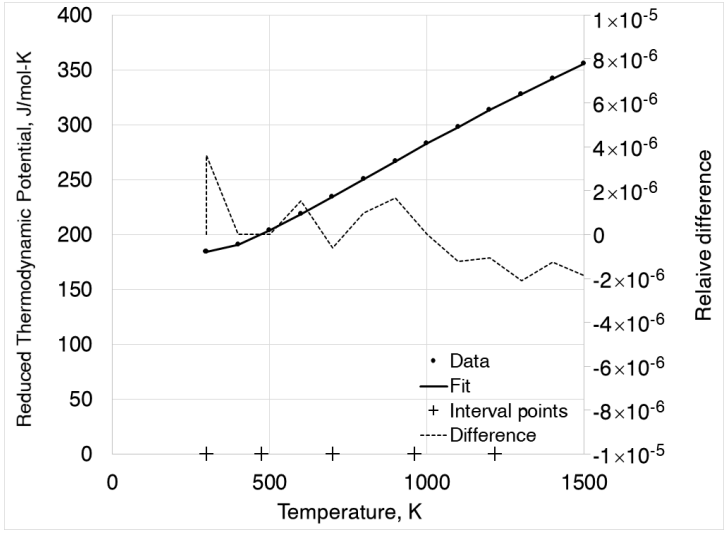
Pentaborane (B_5H_9) – B5H9(l)

MW = 63.12 g/mol – $\Delta_f H^o$(298.15 K) = 42.844 kJ/mol – S^o(298.15 K) = 184.33 J/mol-K – p^o = 0.1 MPa



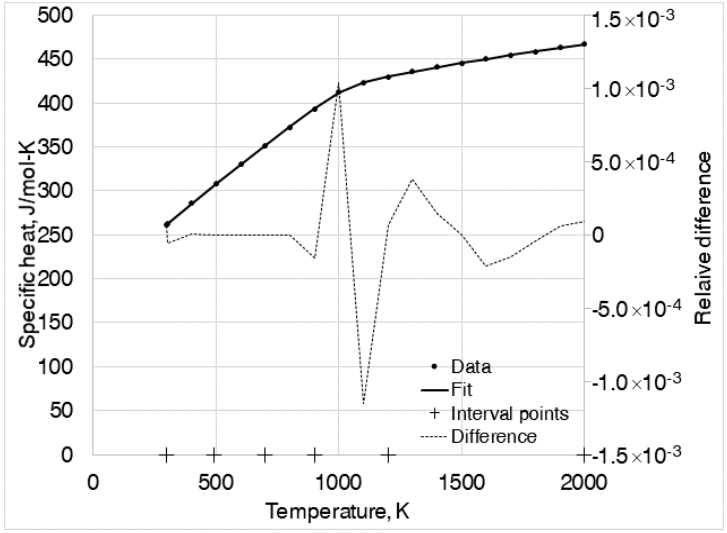


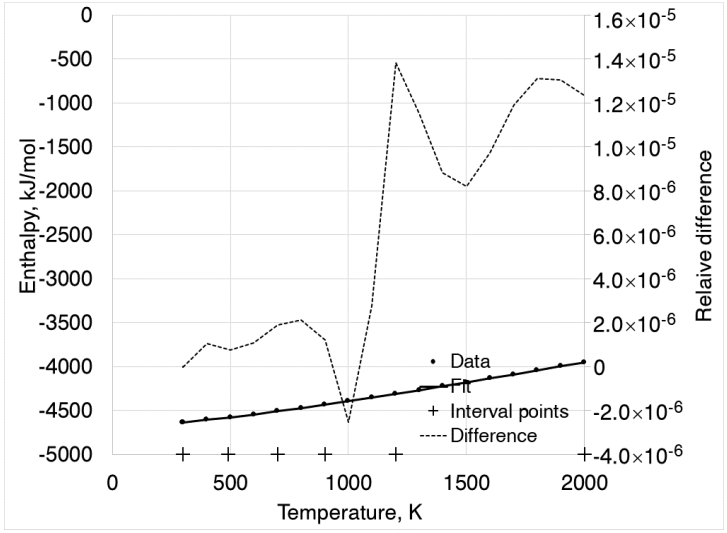


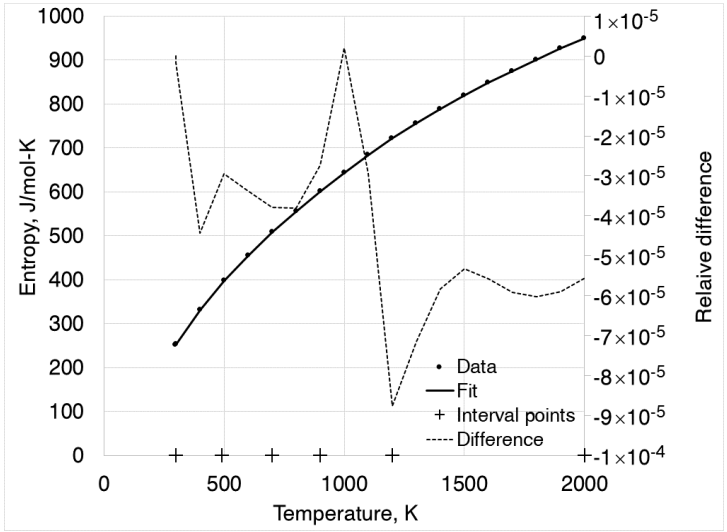


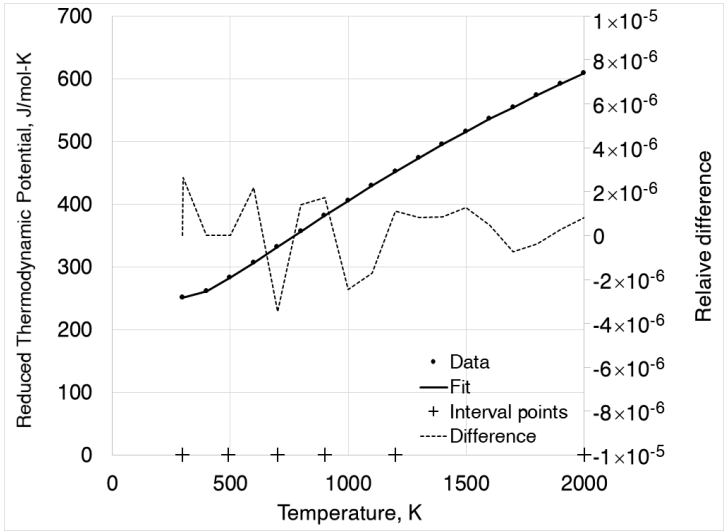
Potassium Borate ($K_2B_6O_{10}$) – B6K2O10(cr)

MW = 303.05 g/mol – $\Delta_f H^o$(298.15 K) = -4633.529 kJ/mol – S^o(298.15 K) = 251.04 J/mol-K – p^o = 0.1 MPa



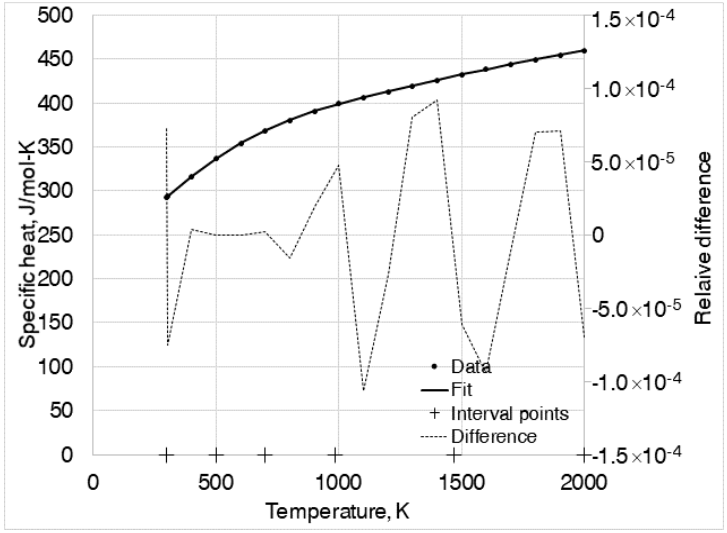


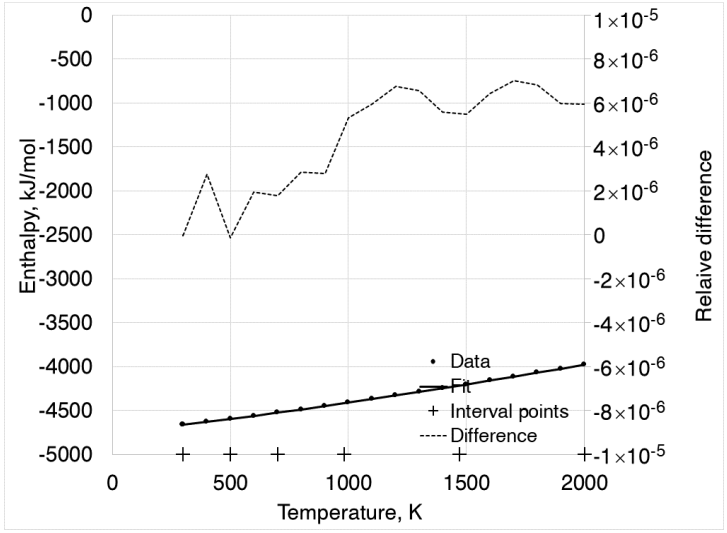


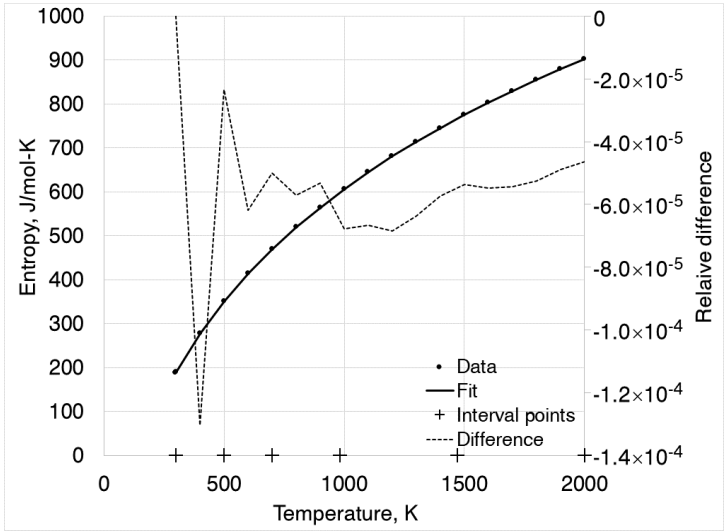


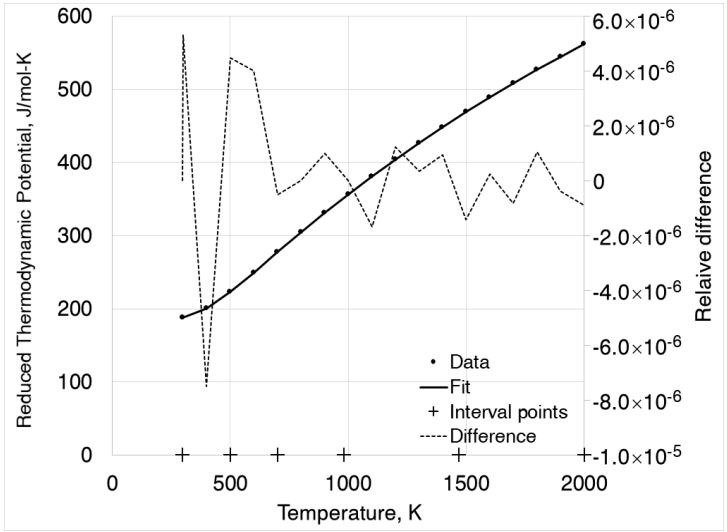
Lithium Borate ($Li_2B_6O_{10}$) – B6Li2O10(cr)

MW = 238.74 g/mol – $\Delta_f H^o$(298.15 K) = -4659.888 kJ/mol – S^o(298.15 K) = 188.28 J/mol-K – p^o = 0.1 MPa



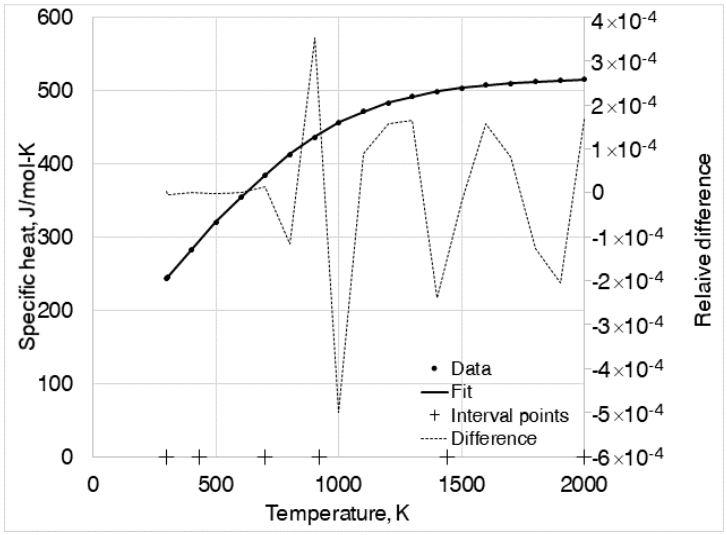


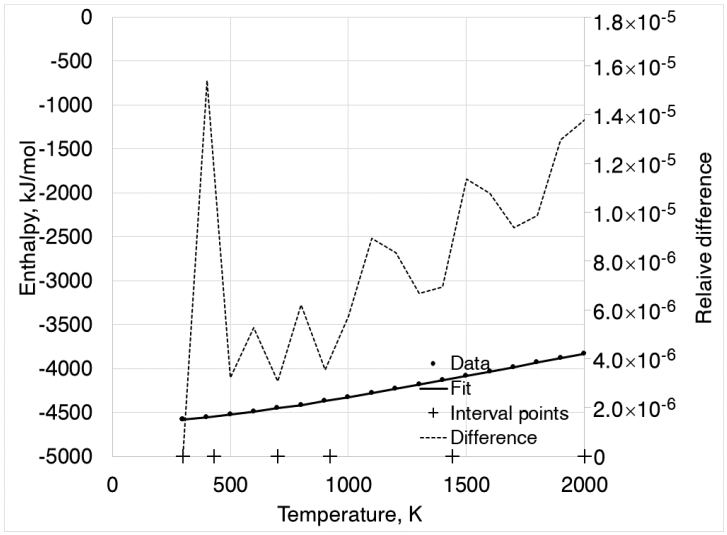


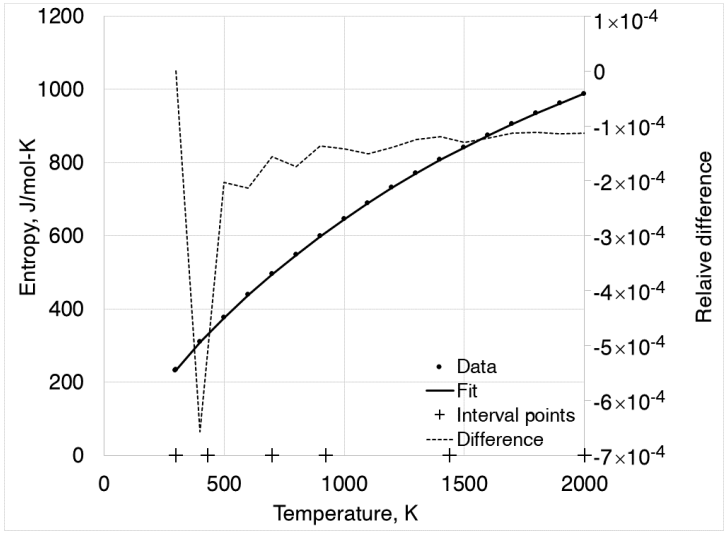


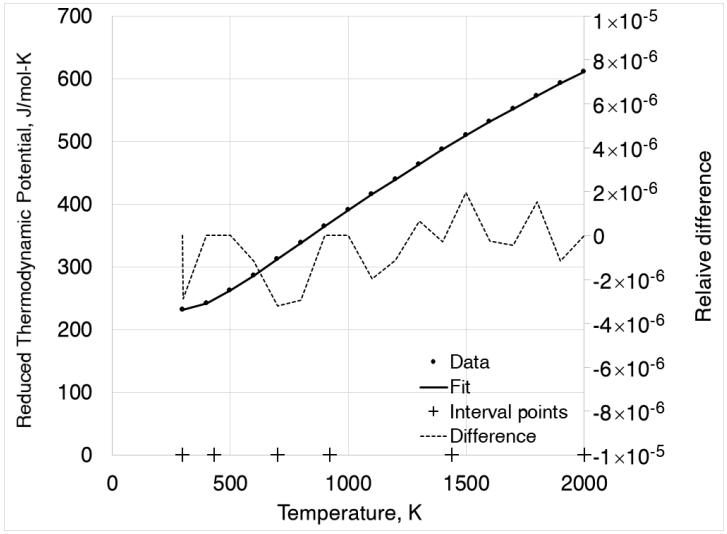
Sodium Borate ($Na_2B_6O_{10}$) – B6Na2O10(cr)

MW = 270.83 g/mol – $\Delta_f H^o$(298.15 K) = -4580.476 kJ/mol – S^o(298.15 K) = 232.212 J/mol-K – p^o = 0.1 MPa



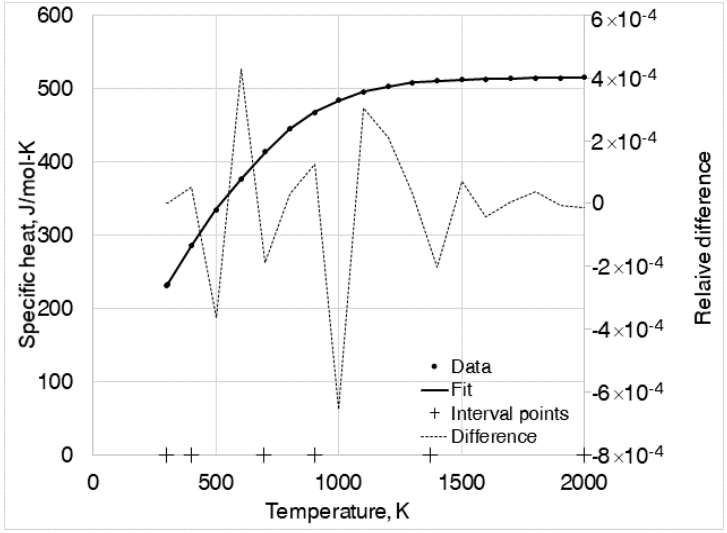


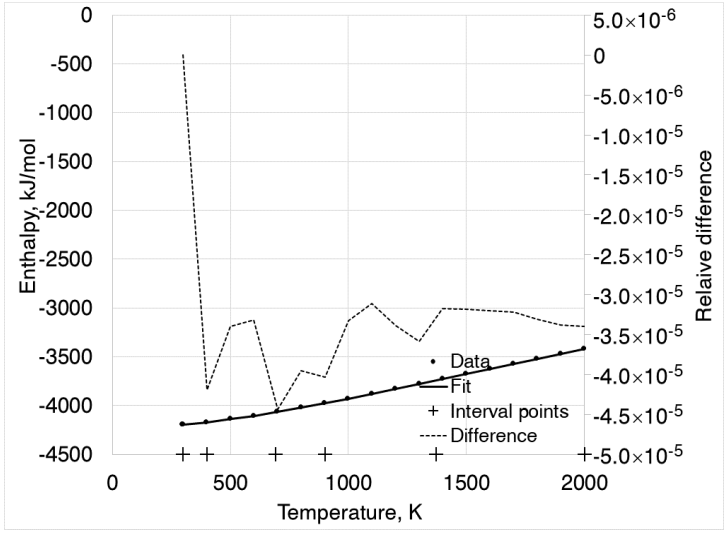


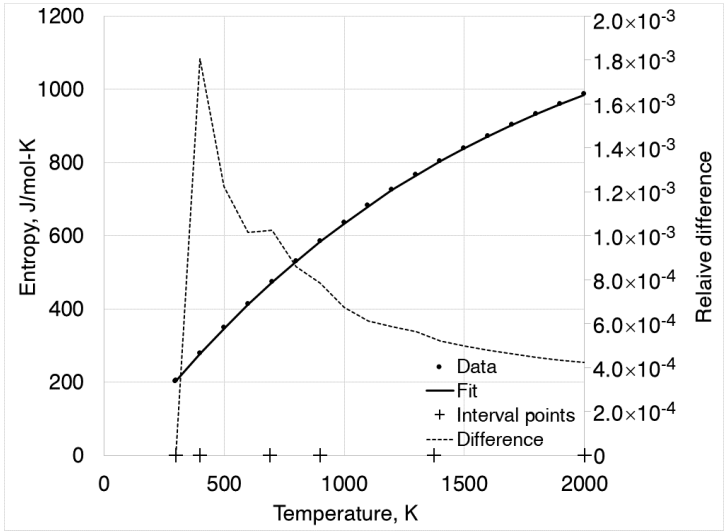


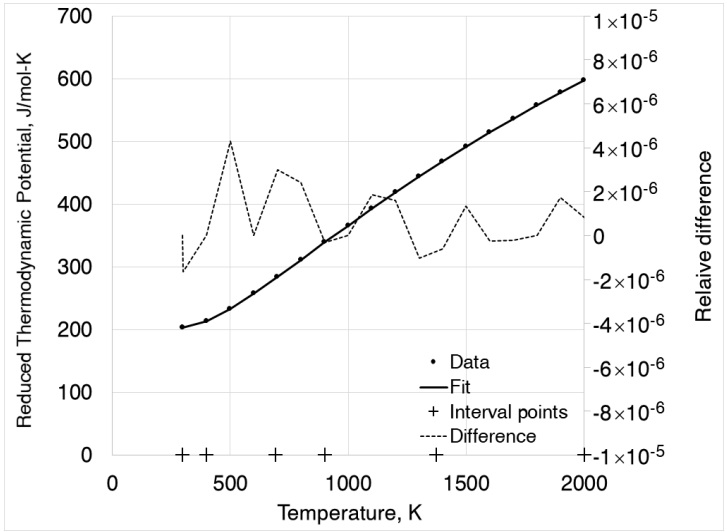
Lead Borate (PbB_6O_{10}) – B6O10Pb1(cr)

MW = 432.1 g/mol – $\Delta_f H^o$(298.15 K) = -4196.552 kJ/mol – S^o(298.15 K) = 203.342 J/mol-K – p^o = 0.1 MPa



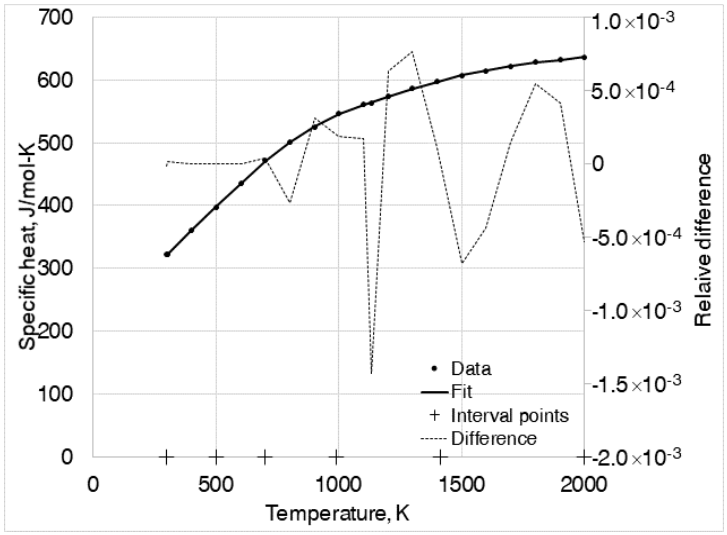


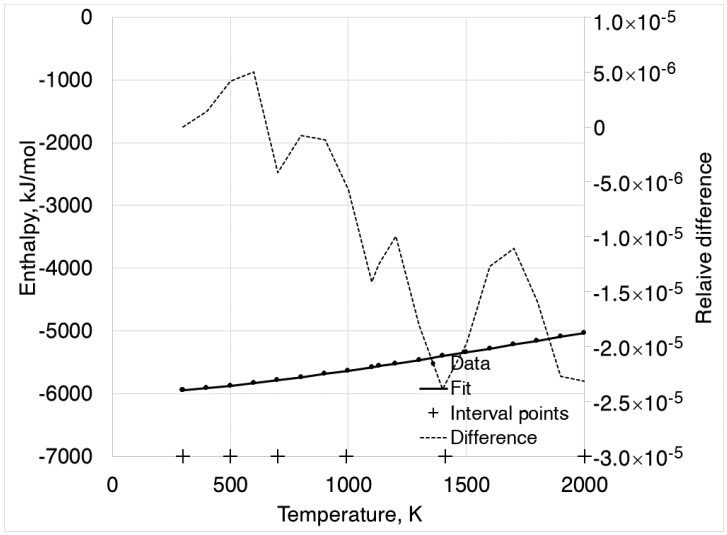


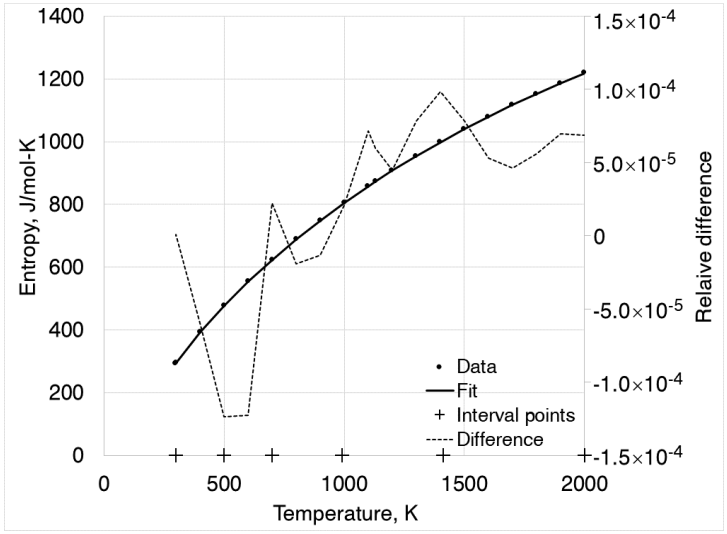


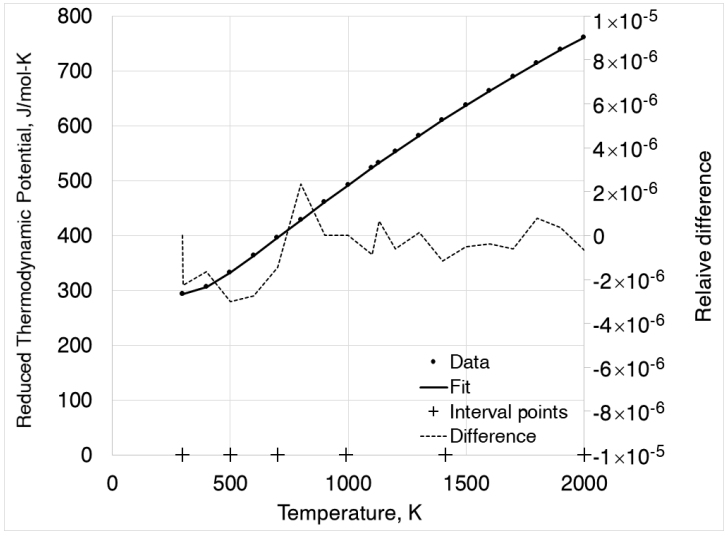
Potassium Borate ($K_2B_8O_{13}$) – B8K2O13(cr)

MW = 372.67 g/mol – $\Delta_f H^o$(298.15 K) = -5945.129 kJ/mol – S^o(298.15 K) = 293.717 J/mol-K – p^o = 0.1 MPa



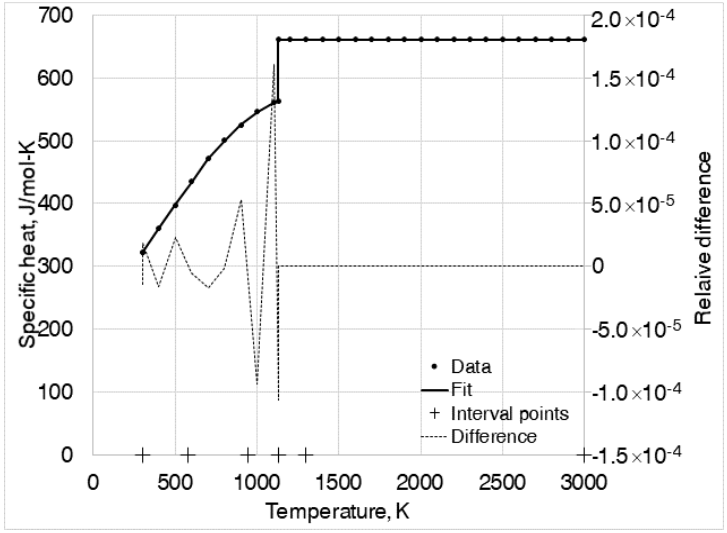


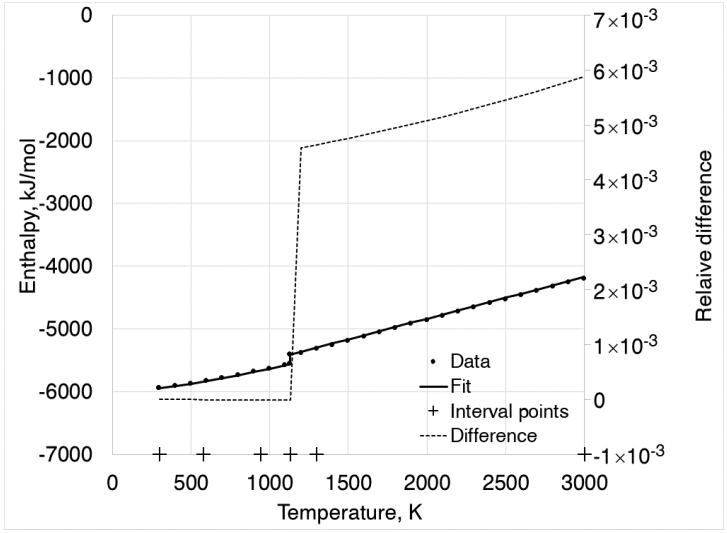


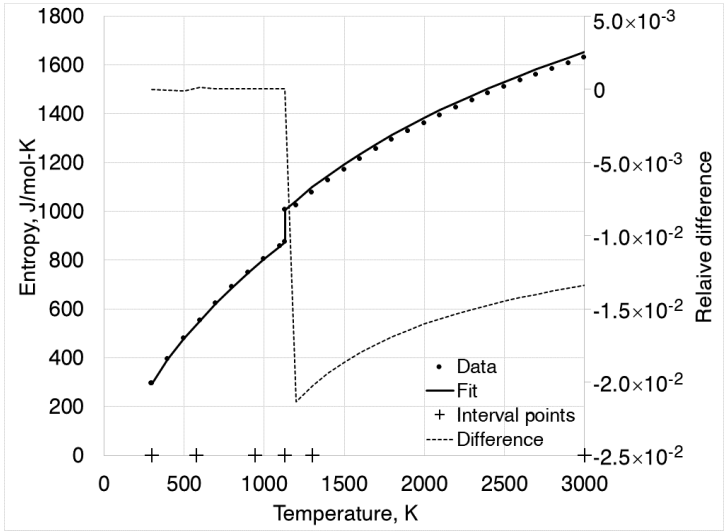


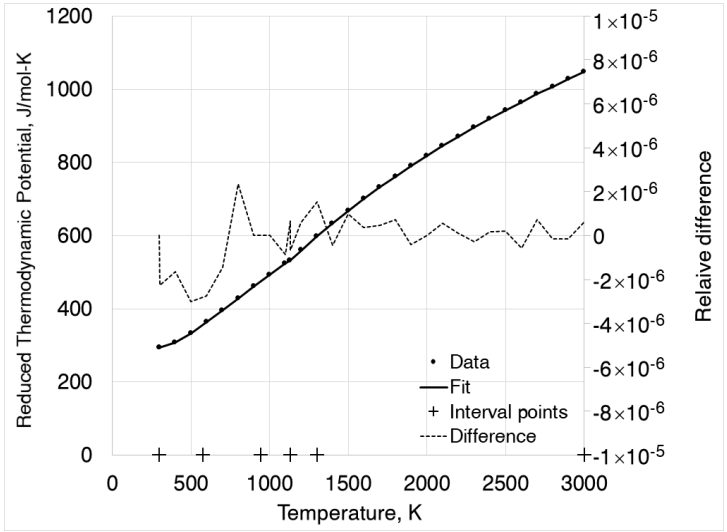
Potassium Borate ($K_2B_8O_{13}$) – B8K2O13(cr,l)

MW = 372.67 g/mol – $\Delta_f H^o$(298.15 K) = -5945.129 kJ/mol – S^o(298.15 K) = 293.717 J/mol-K – p^o = 0.1 MPa



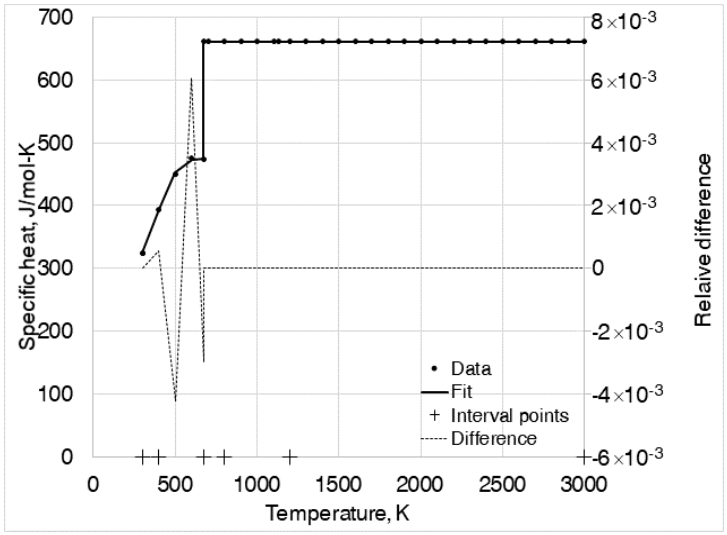


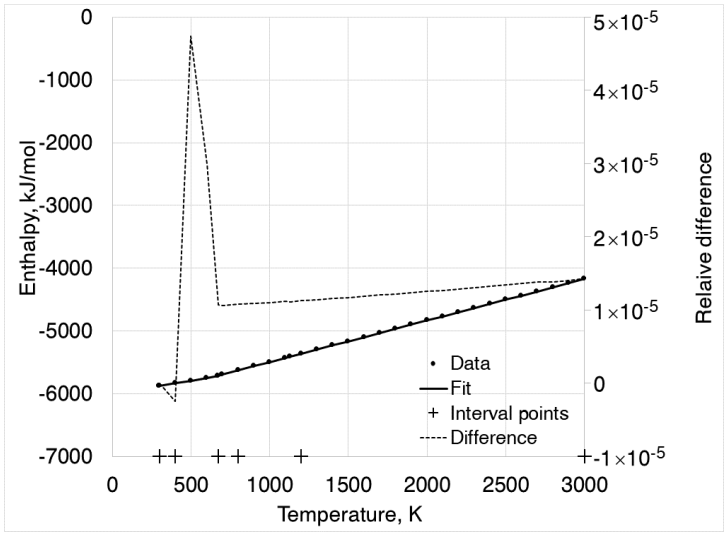


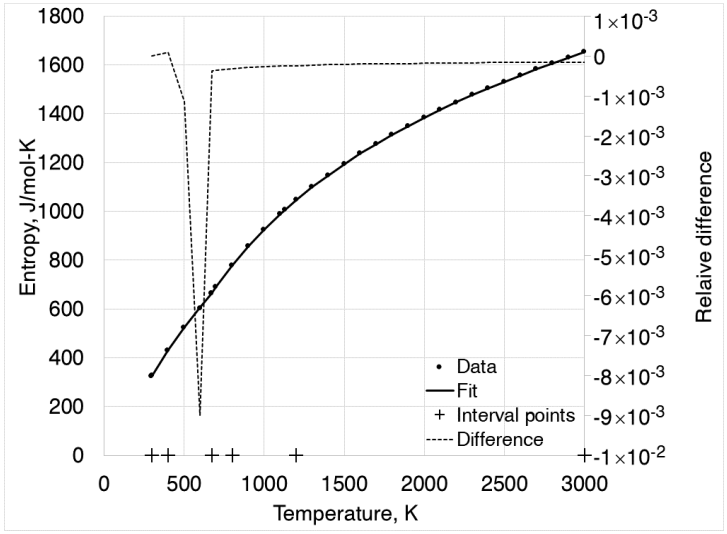


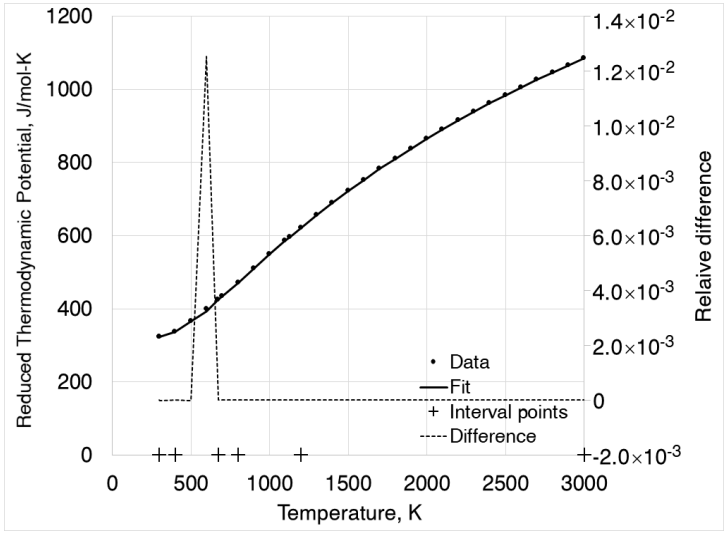
Potassium Borate ($K_2B_8O_{13}$) – B8K2O13(l)

MW = 372.67 g/mol – $\Delta_f H^o$(298.15 K) = -5872.537 kJ/mol – S^o(298.15 K) = 324.085 J/mol-K – p^o = 0.1 MPa



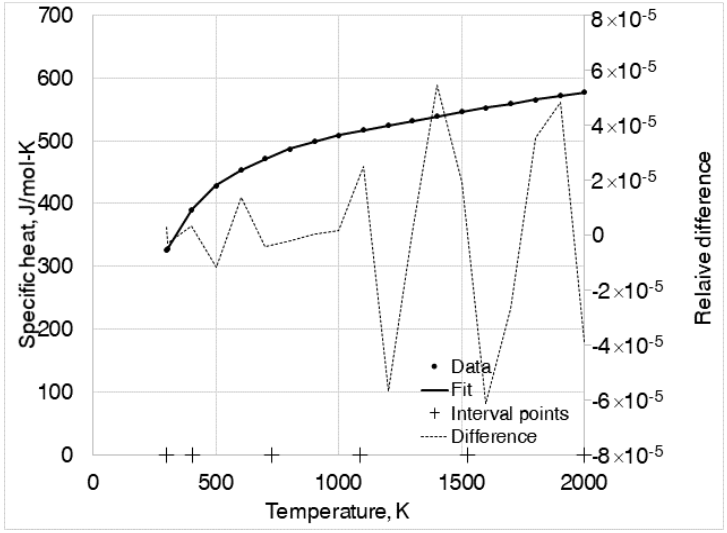


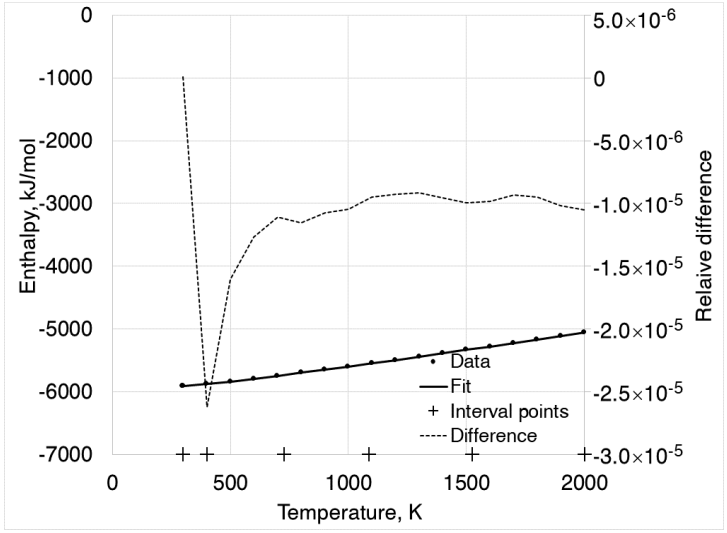


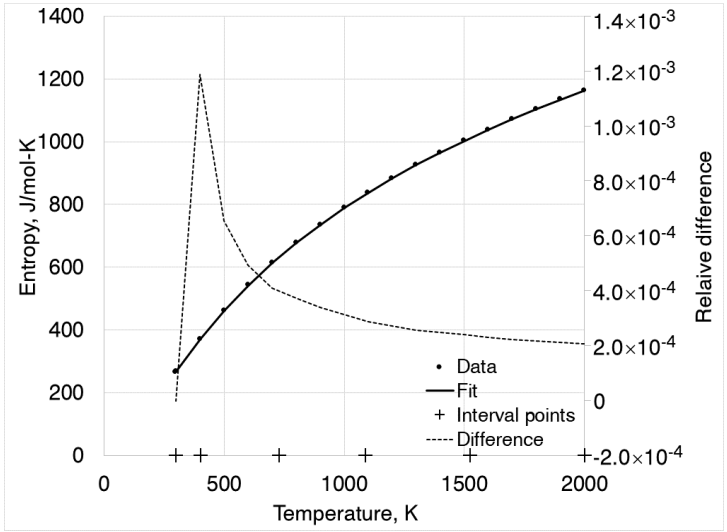


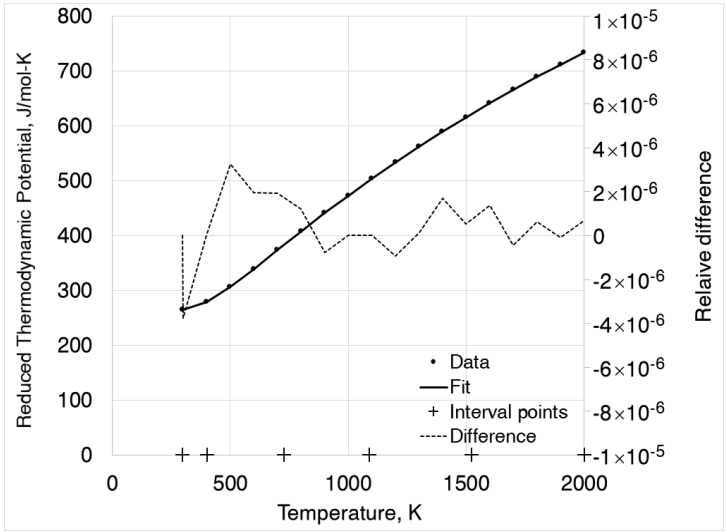
Lithium Borate ($Li_2B_8O_{13}$) – B8Li2O13(cr)

MW = 308.35 g/mol – $\Delta_f H^o$(298.15 K) = -5914.377 kJ/mol – S^o(298.15 K) = 265.266 J/mol-K – p^o = 0.1 MPa



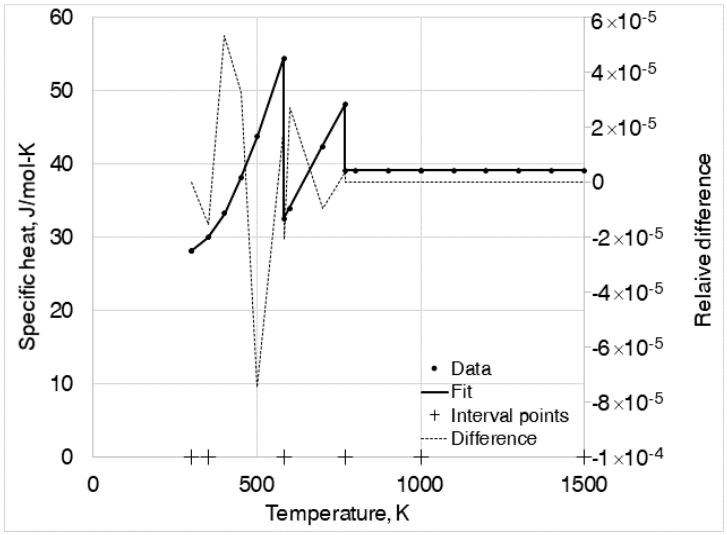


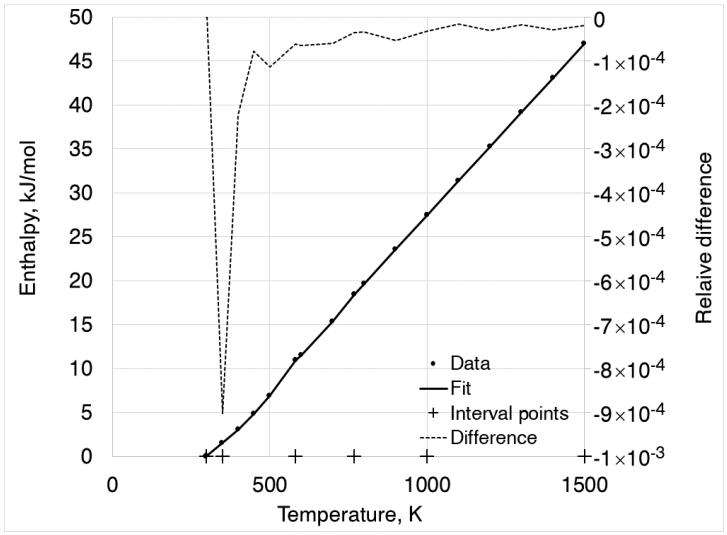


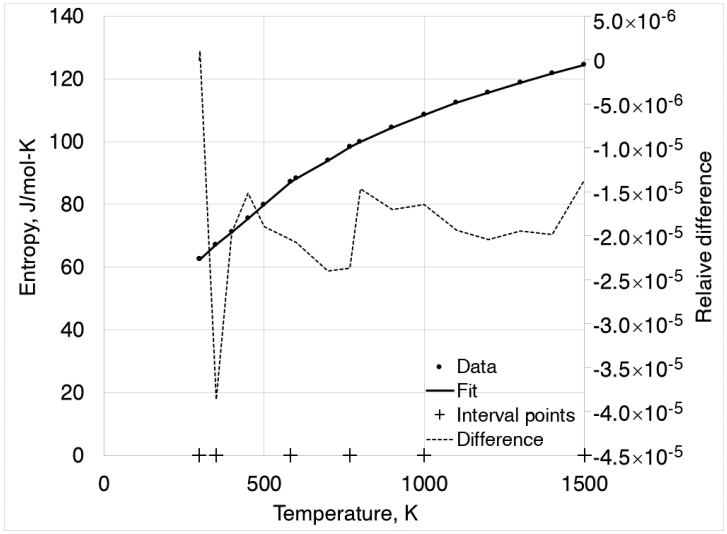


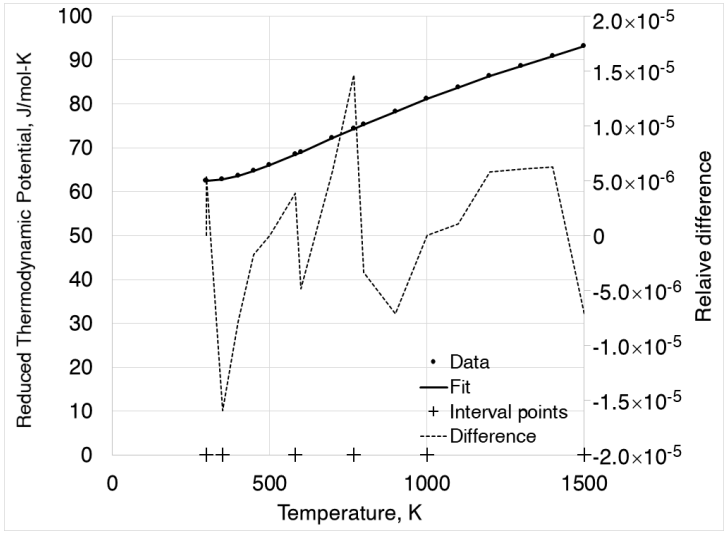
Barium (Ba) – Ba1(cr)

MW = 137.33 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 62.475 J/mol-K – p^o = 0.1 MPa



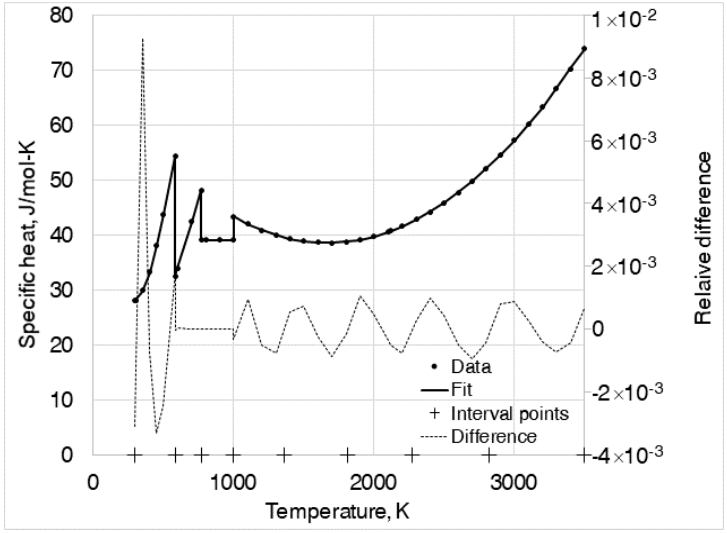


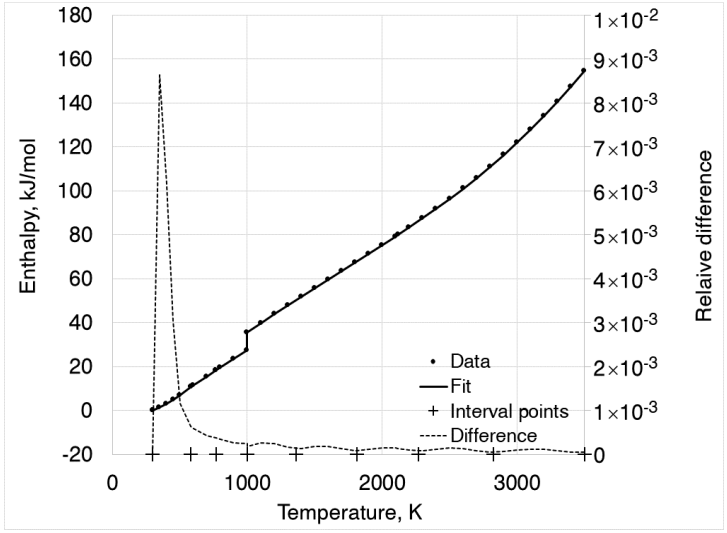


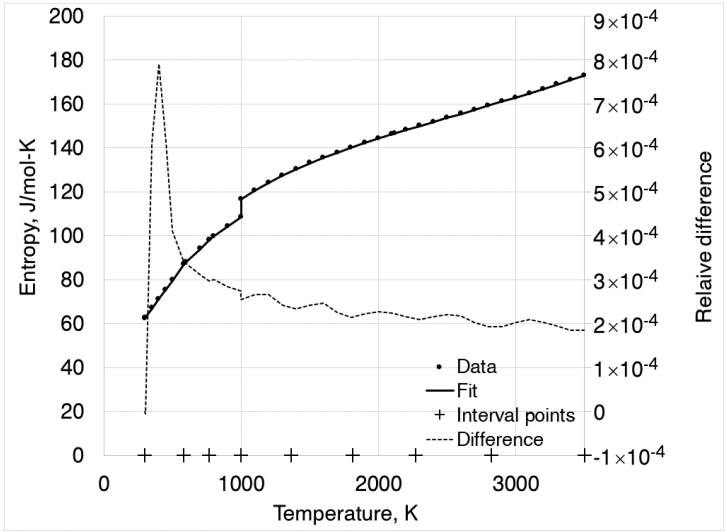


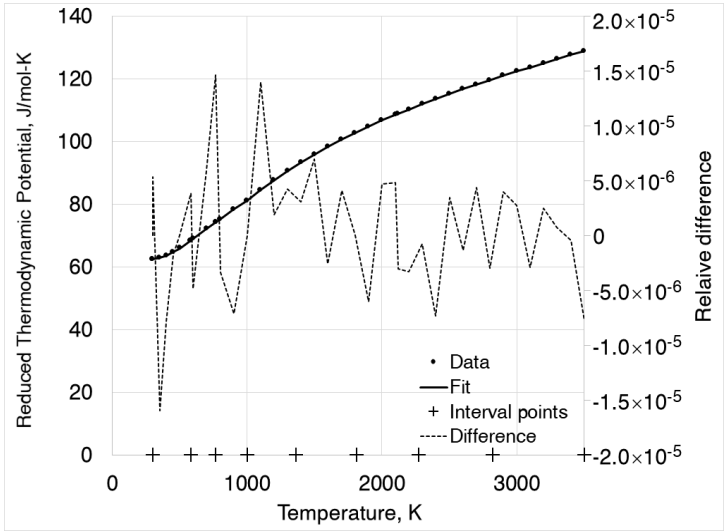
Barium (Ba) – Ba1(cr,l)

MW = 137.33 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 62.475 J/mol-K – p^o = 0.1 MPa



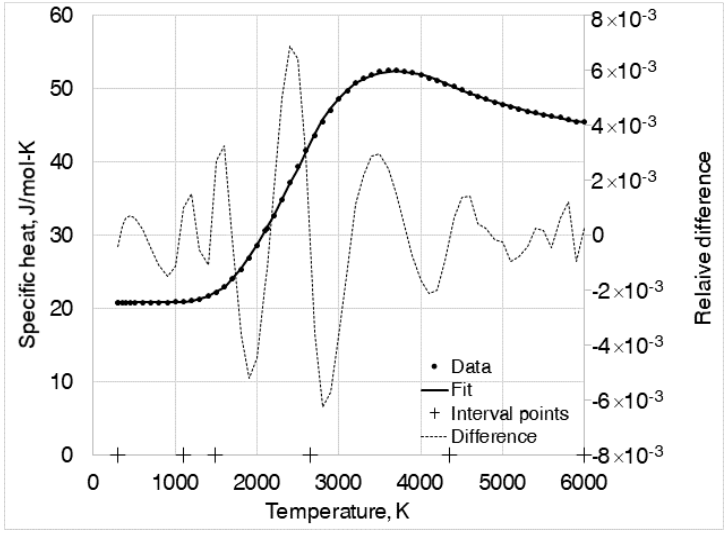


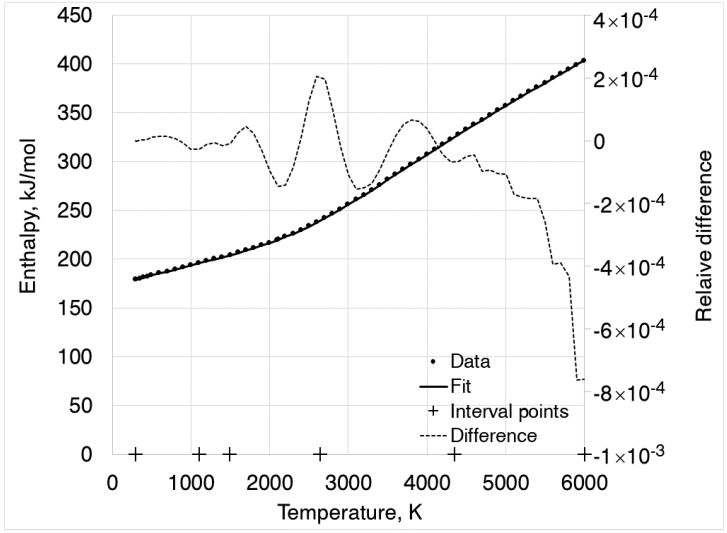


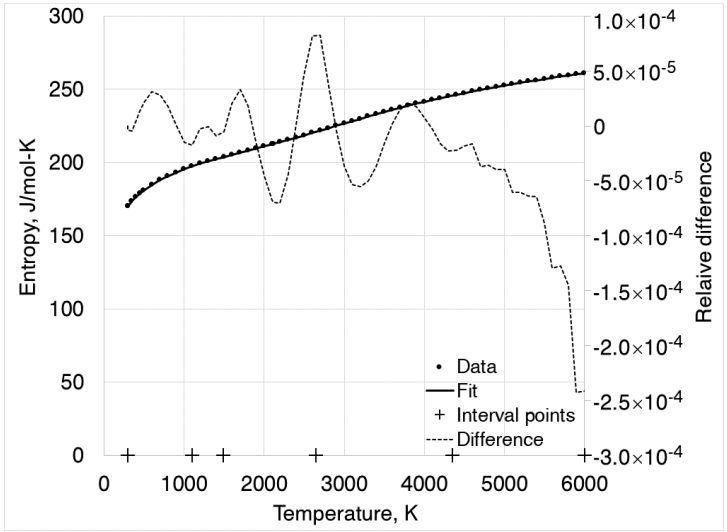


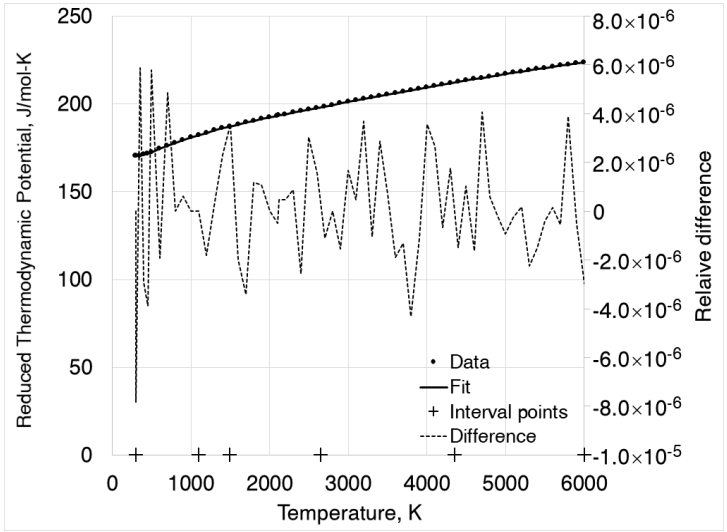
Barium (Ba) – Ba1(g)

MW = 137.33 g/mol – $\Delta_f H^o$(298.15 K) = 179.075 kJ/mol – S^o(298.15 K) = 170.245 J/mol-K – p^o = 0.1 MPa



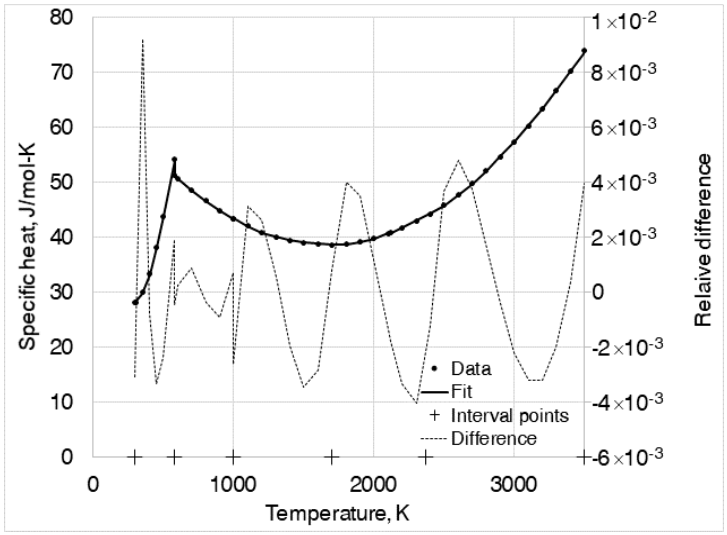


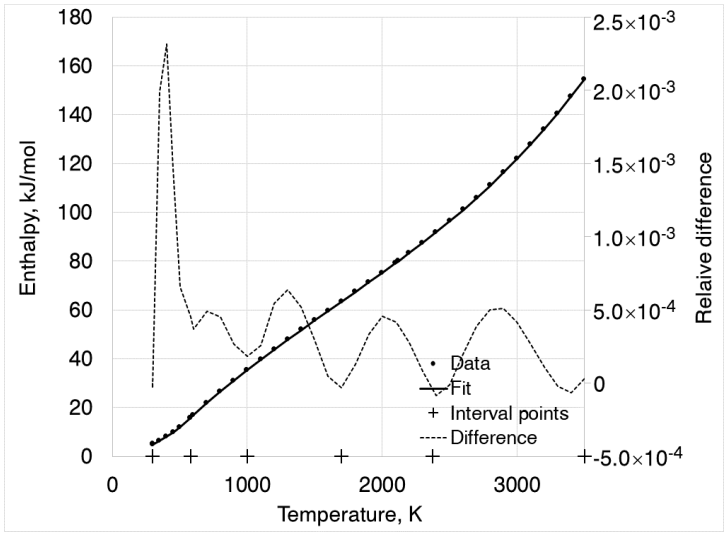


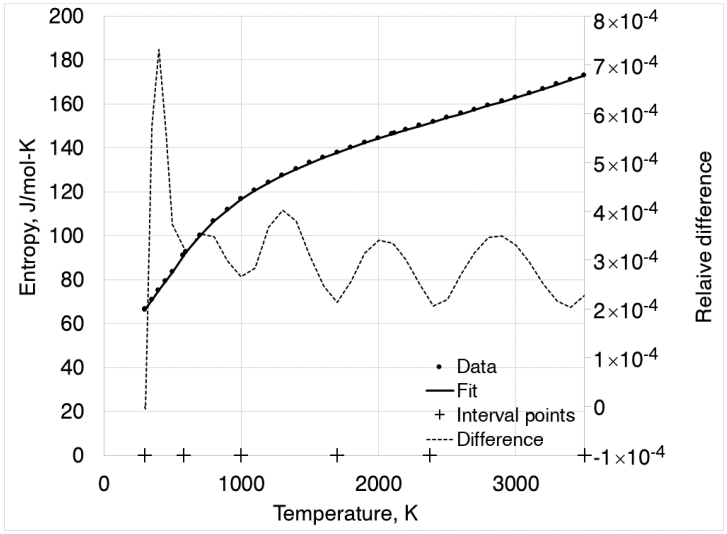


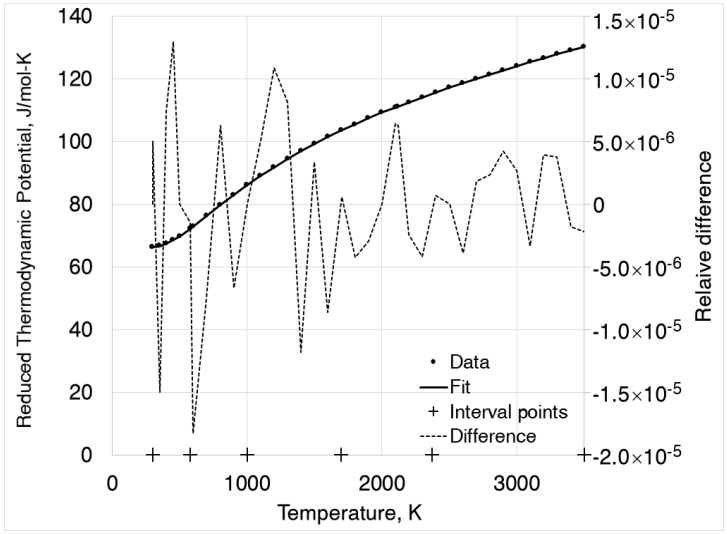
Barium (Ba) – Ba1(l)

MW = 137.33 g/mol – $\Delta_f H^o$(298.15 K) = 4.985 kJ/mol – S^o(298.15 K) = 66.328 J/mol-K – p^o = 0.1 MPa



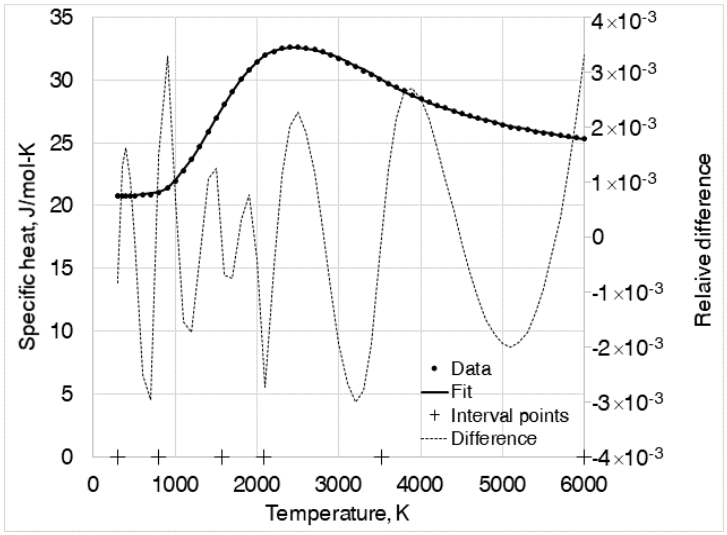


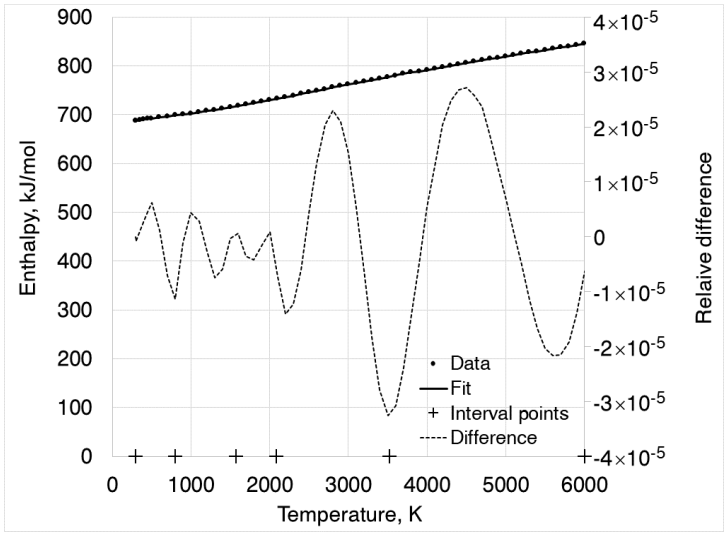


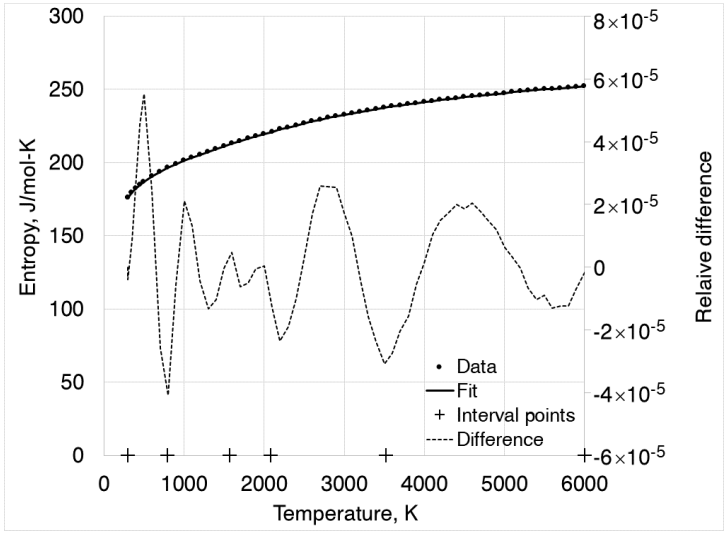


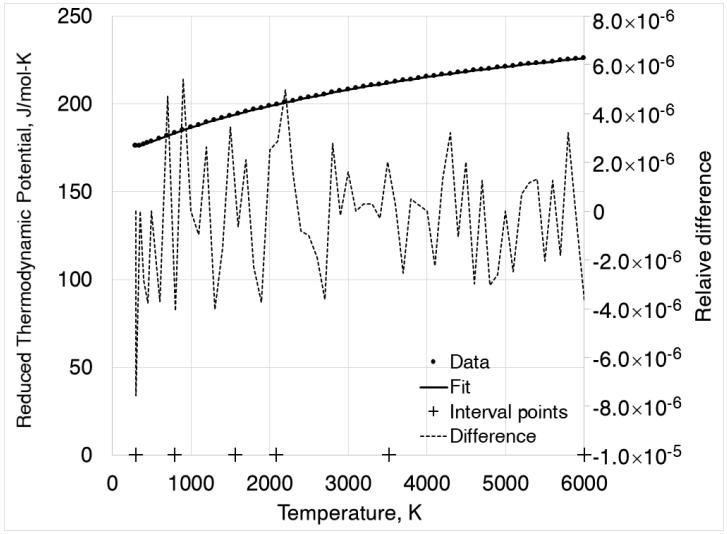
Barium, Ion (Ba^+) – Ba1+(g)

MW = 137.33 g/mol – $\Delta_f H^o$(298.15 K) = 688.126 kJ/mol – S^o(298.15 K) = 176.008 J/mol-K – p^o = 0.1 MPa



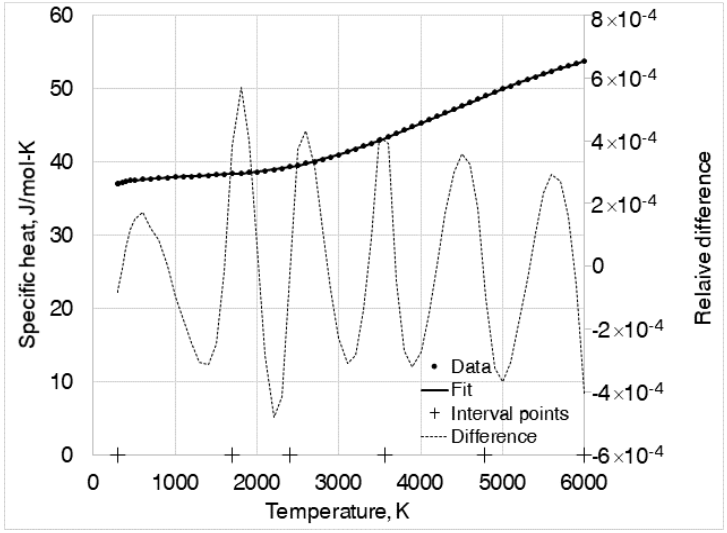


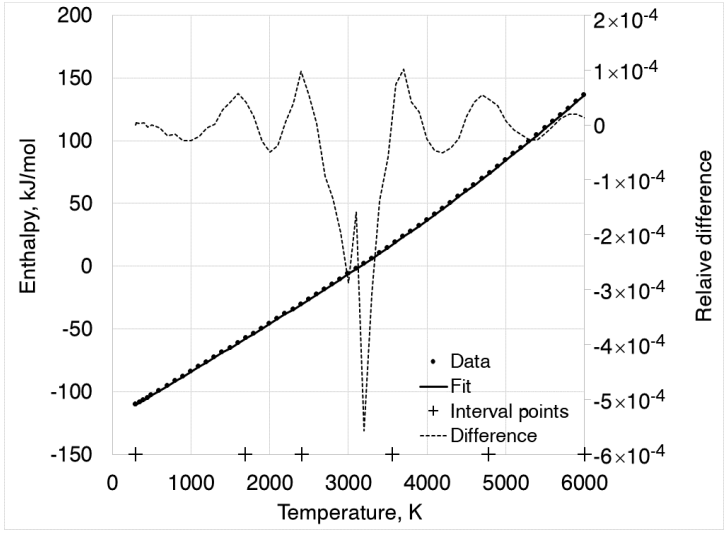


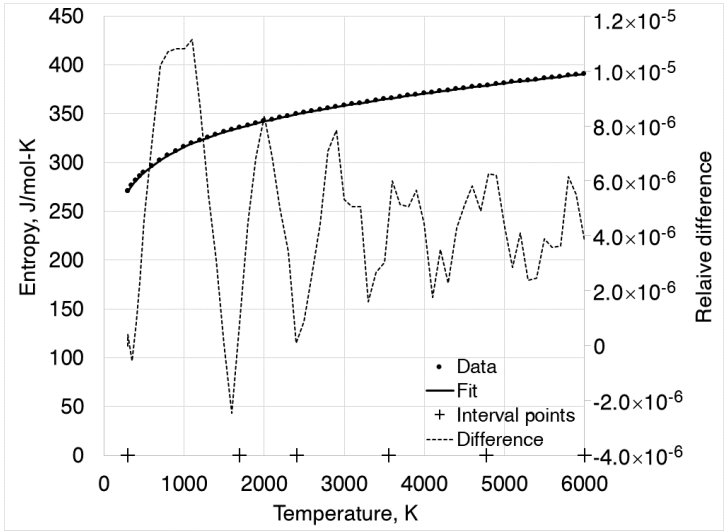


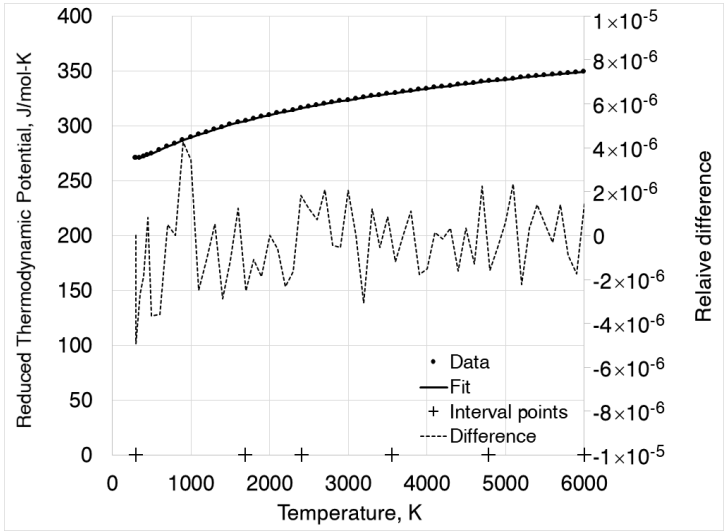
Barium Bromide (BaBr) – Ba1Br1(g)

MW = 217.23 g/mol – $\Delta_f H^o$(298.15 K) = -110.587 kJ/mol – S^o(298.15 K) = 270.395 J/mol-K – p^o = 0.1 MPa



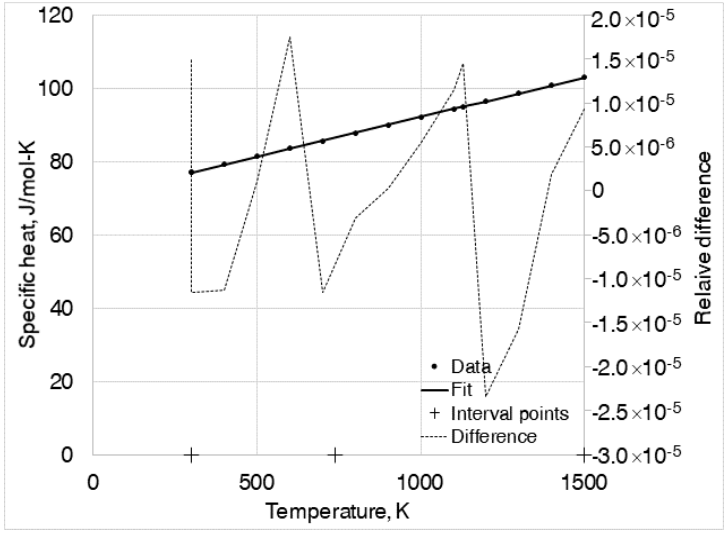


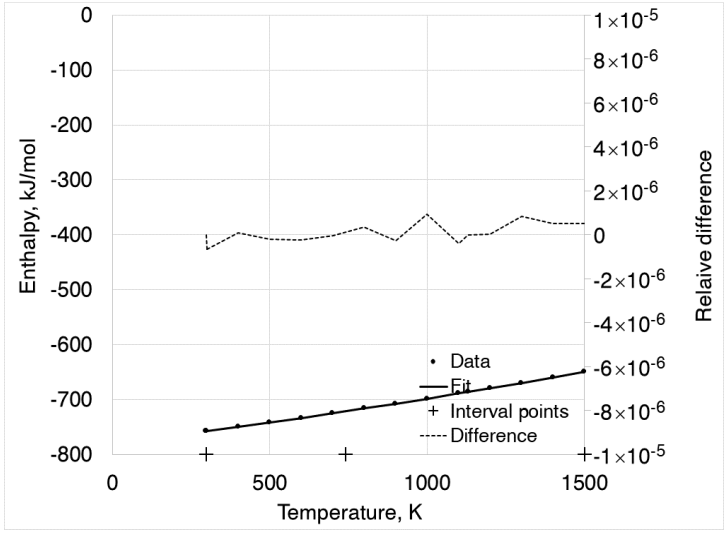


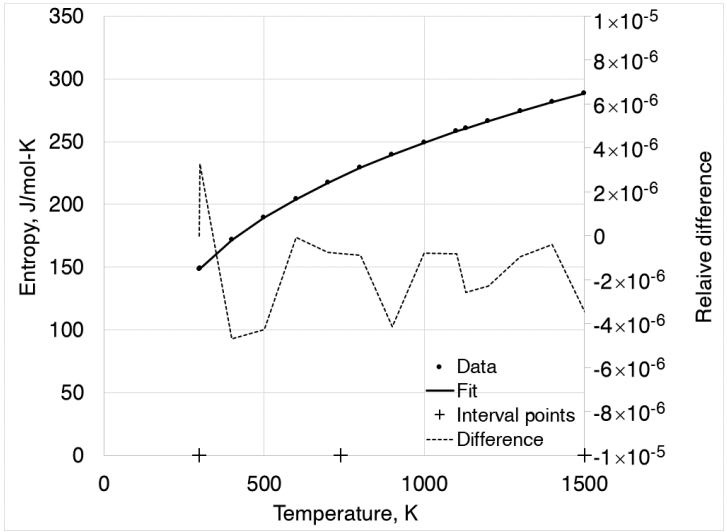


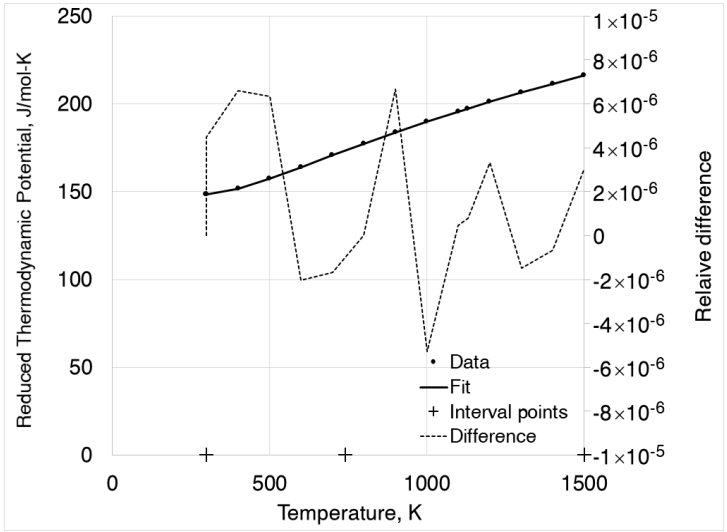
Barium Bromide ($BaBr_2$) – Ba1Br2(cr)

MW = 297.14 g/mol – $\Delta_f H^o$(298.15 K) = -757.722 kJ/mol – S^o(298.15 K) = 148.532 J/mol-K – p^o = 0.1 MPa



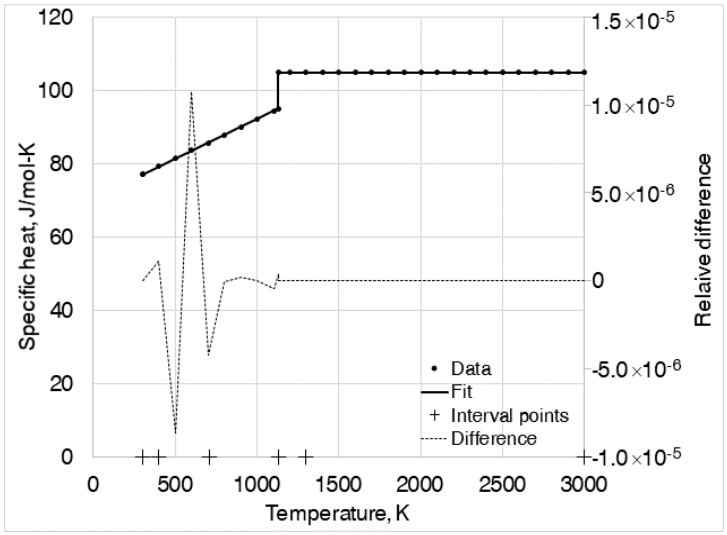


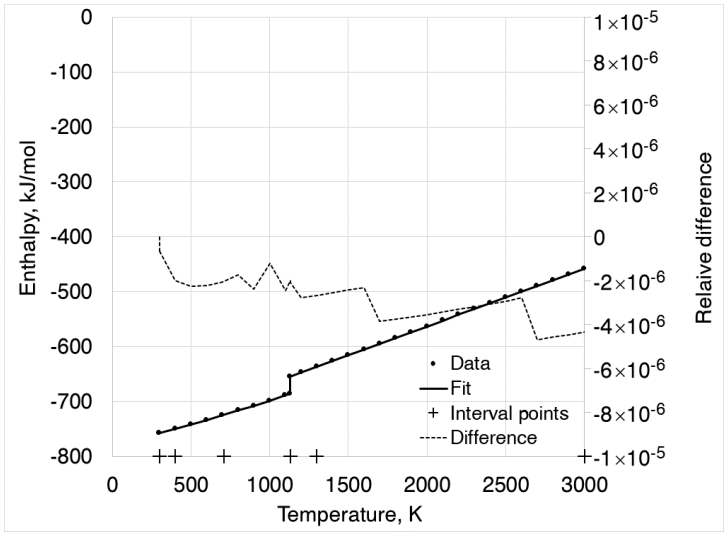


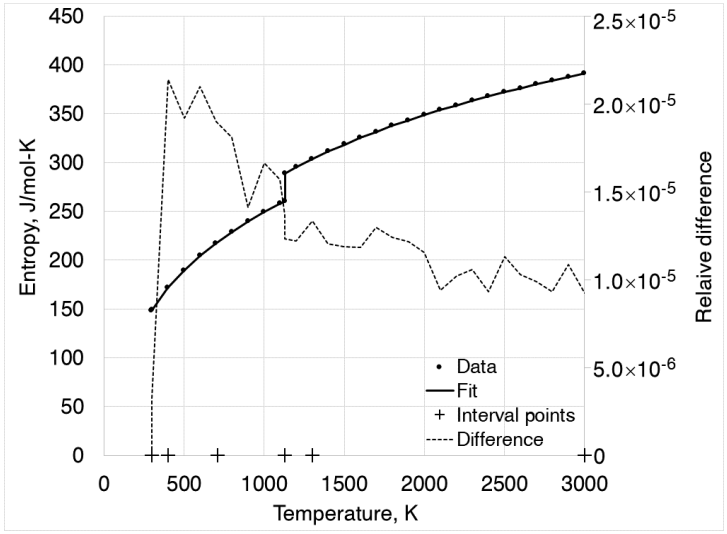


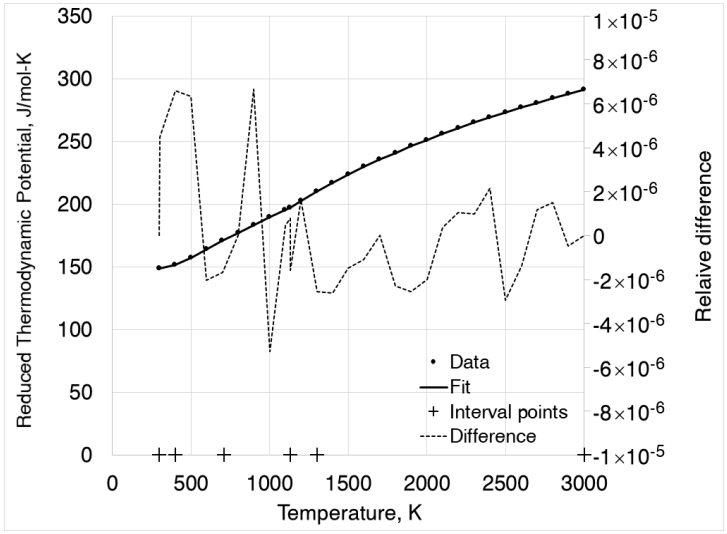
Barium Bromide ($BaBr_2$) – Ba1Br2(cr,l)

MW = 297.14 g/mol – $\Delta_f H^o$(298.15 K) = -757.722 kJ/mol – S^o(298.15 K) = 148.532 J/mol-K – p^o = 0.1 MPa



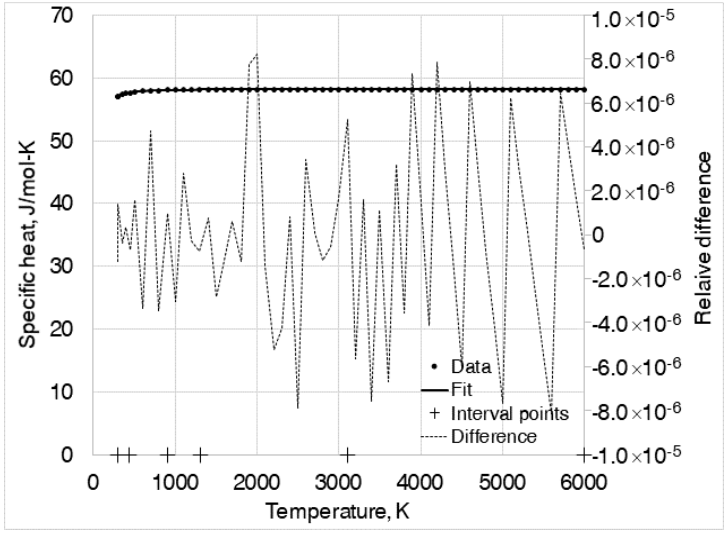


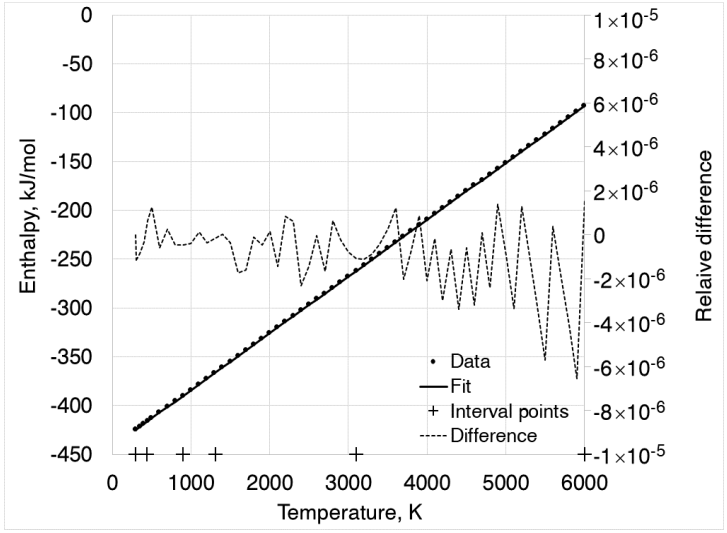


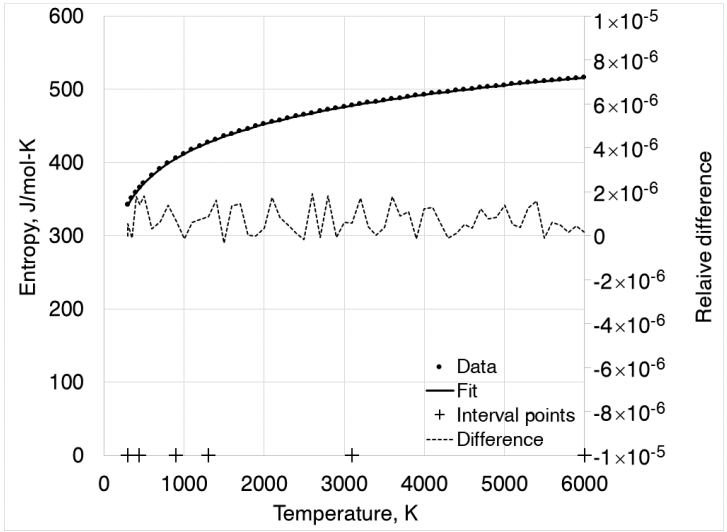


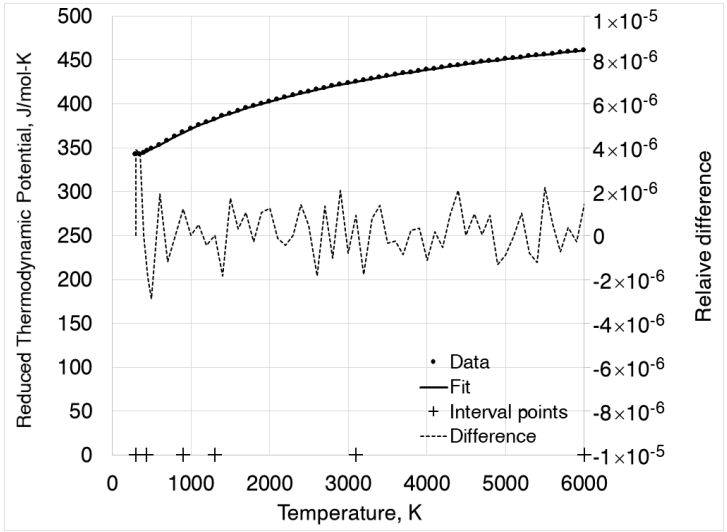
Barium Bromide ($BaBr_2$) – Ba1Br2(g)

MW = 297.14 g/mol – $\Delta_f H^o$(298.15 K) = -424.676 kJ/mol – S^o(298.15 K) = 342.067 J/mol-K – p^o = 0.1 MPa



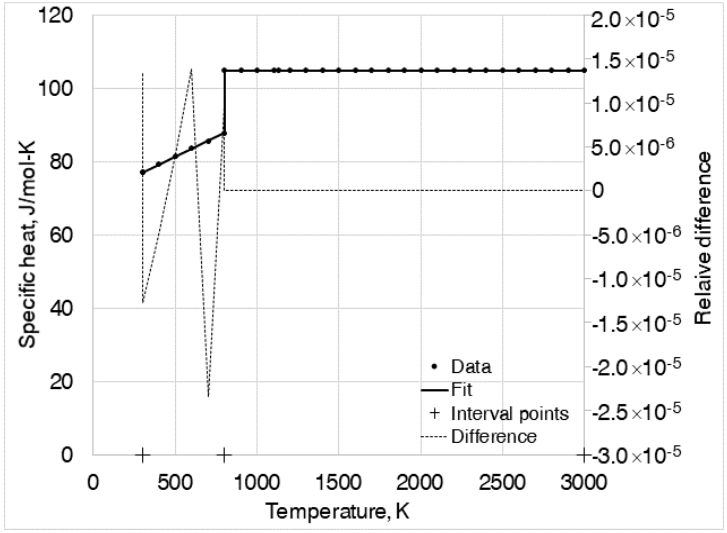


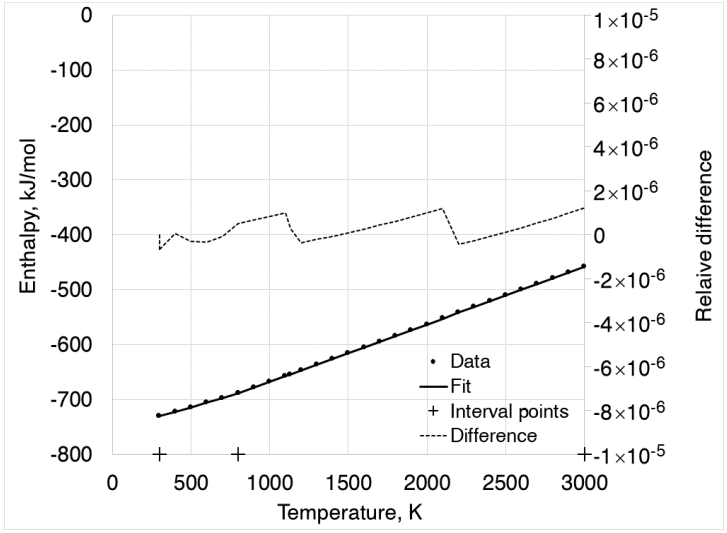


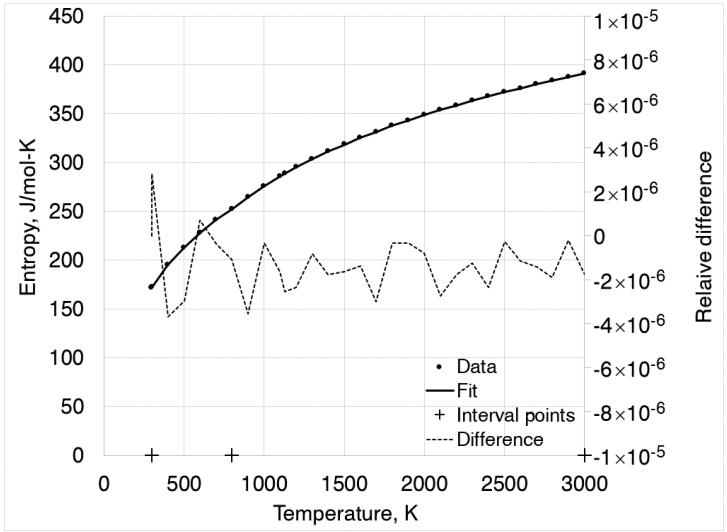


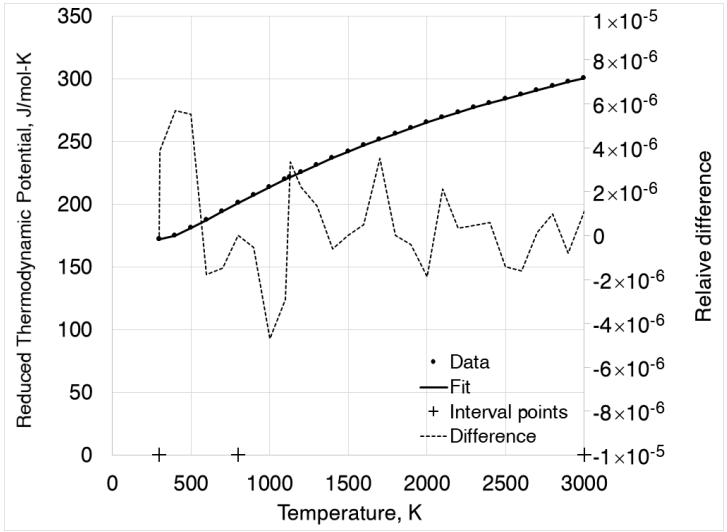
Barium Bromide ($BaBr_2$) – Ba1Br2(l)

MW = 297.14 g/mol – $\Delta_f H^o$(298.15 K) = -730.193 kJ/mol – S^o(298.15 K) = 172.108 J/mol-K – p^o = 0.1 MPa



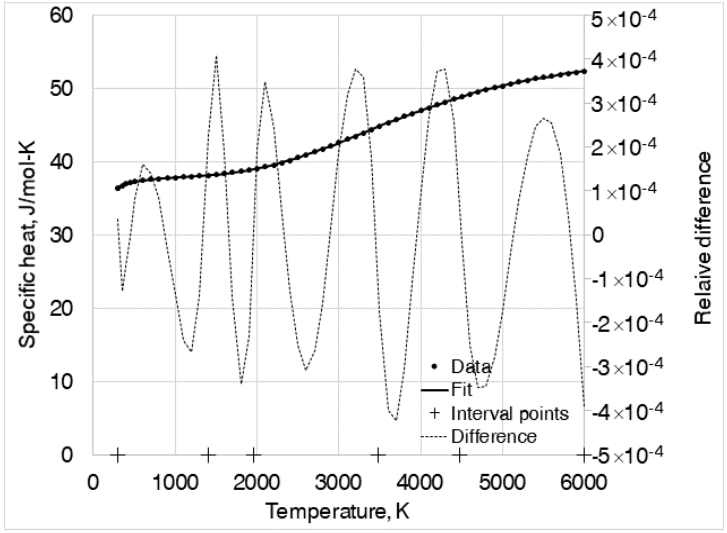


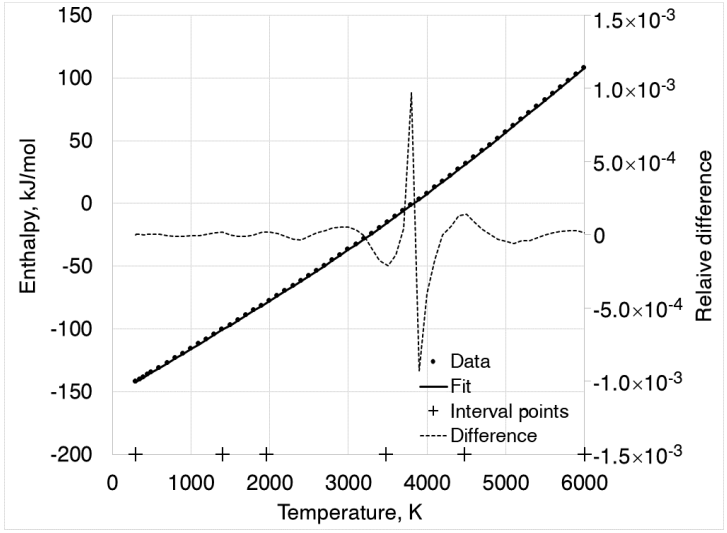


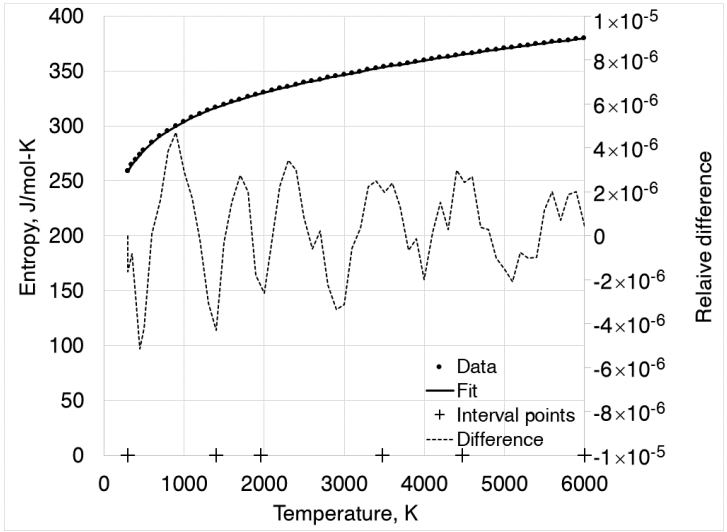


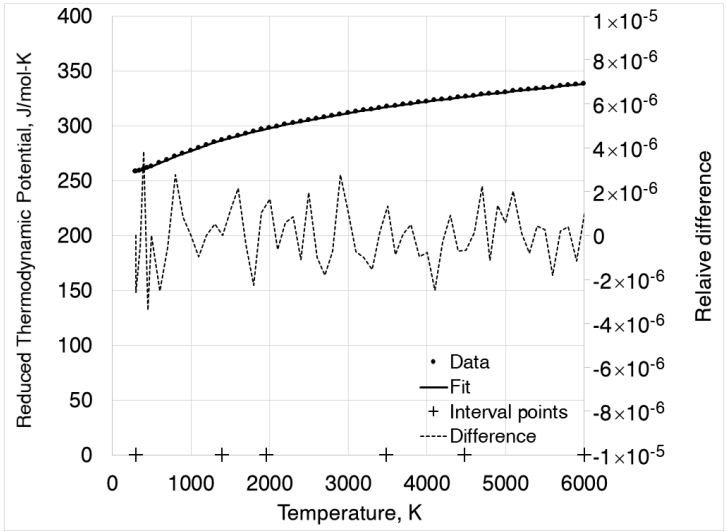
Barium Chloride (BaCl) – Ba1Cl1(g)

MW = 172.78 g/mol – $\Delta_f H^o$(298.15 K) = -142.256 kJ/mol – S^o(298.15 K) = 258.643 J/mol-K – p^o = 0.1 MPa



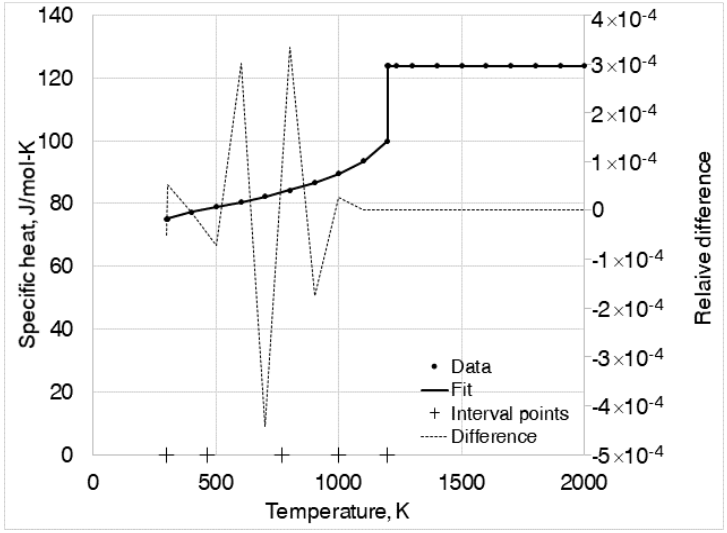


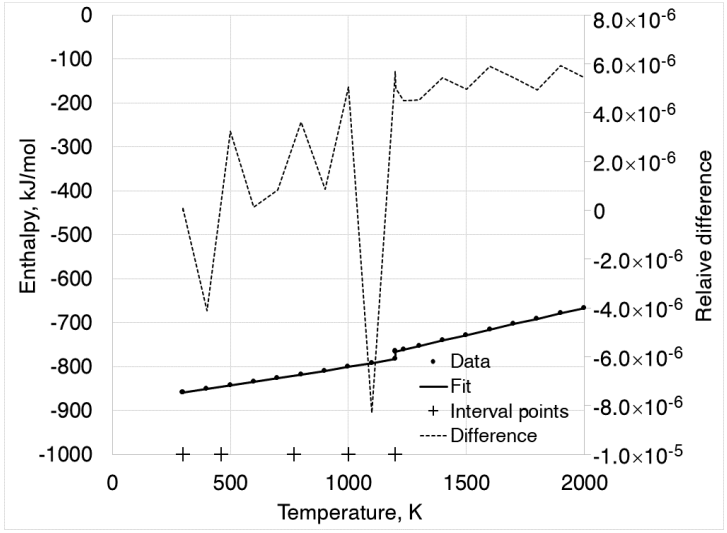


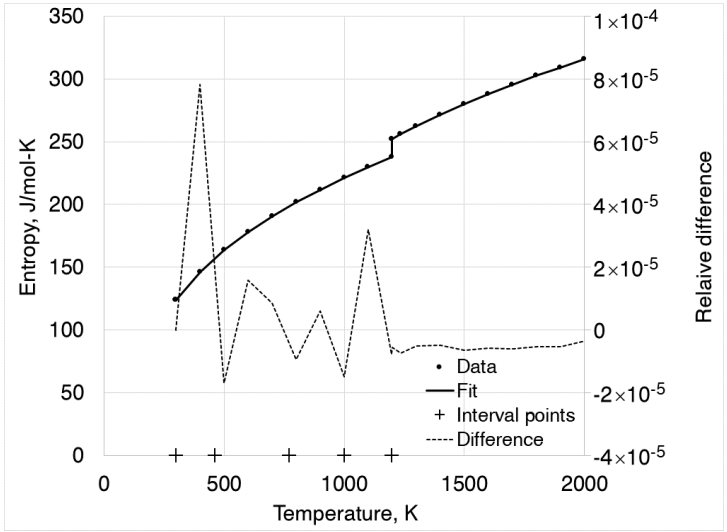


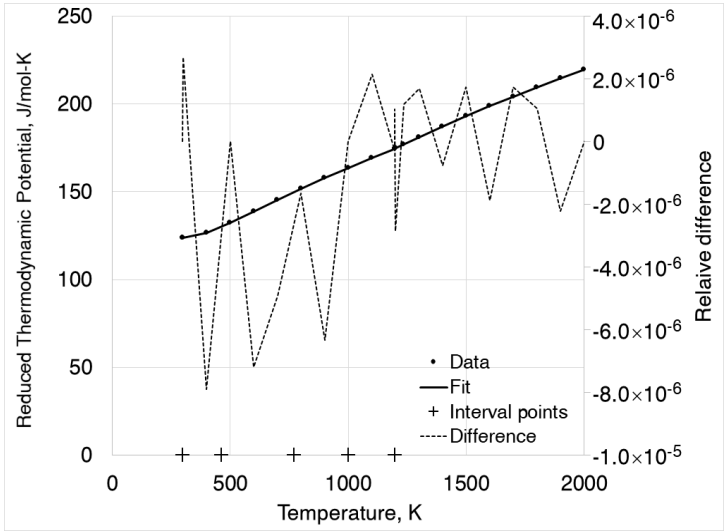
Barium Chloride ($BaCl_2$) – Ba1Cl2(cr)

MW = 208.24 g/mol – $\Delta_f H^o$(298.15 K) = -858.557 kJ/mol – S^o(298.15 K) = 123.666 J/mol-K – p^o = 0.1 MPa



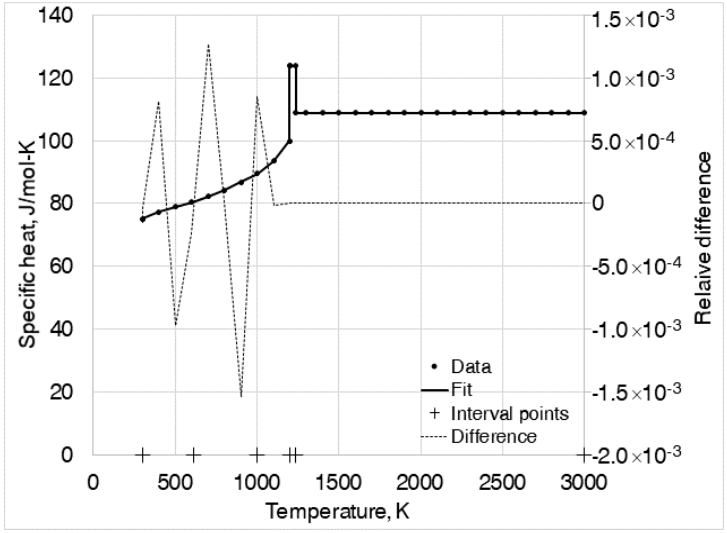


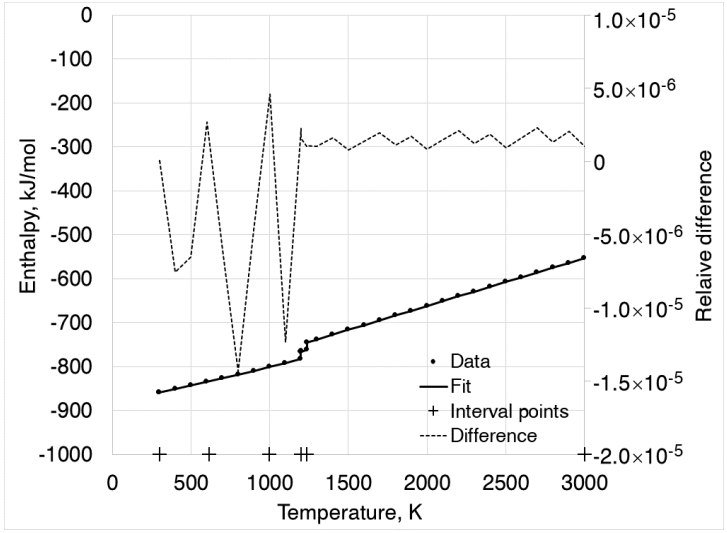


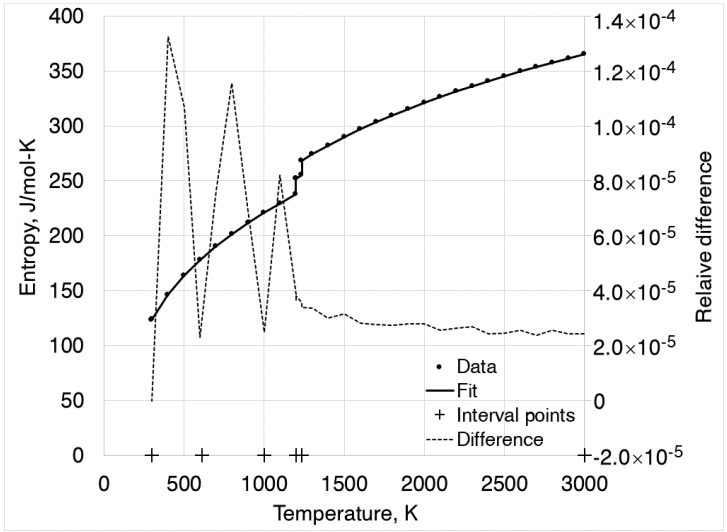


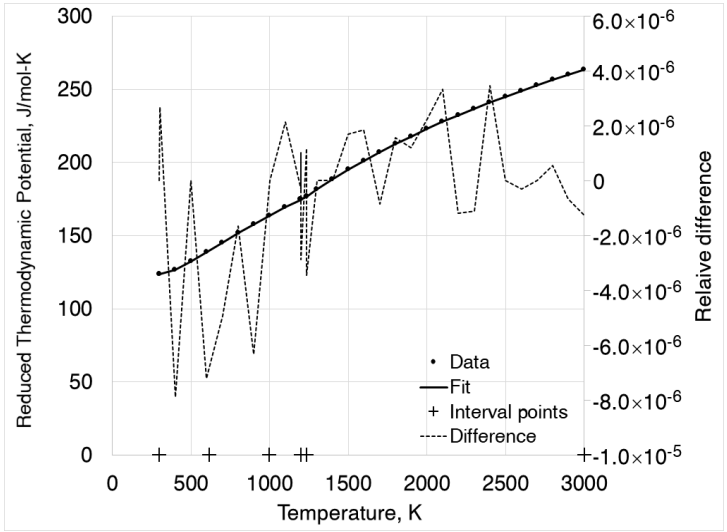
Barium Chloride ($BaCl_2$) – Ba1Cl2(cr,l)

MW = 208.24 g/mol – $\Delta_f H^o$(298.15 K) = -858.557 kJ/mol – S^o(298.15 K) = 123.666 J/mol-K – p^o = 0.1 MPa



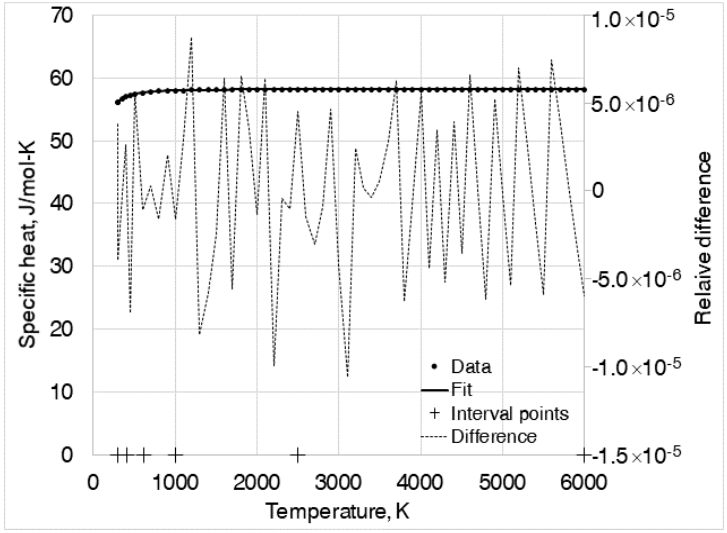


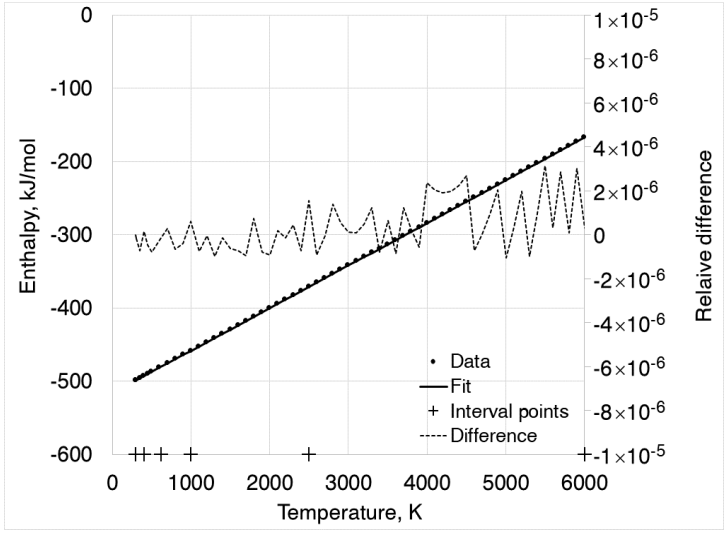


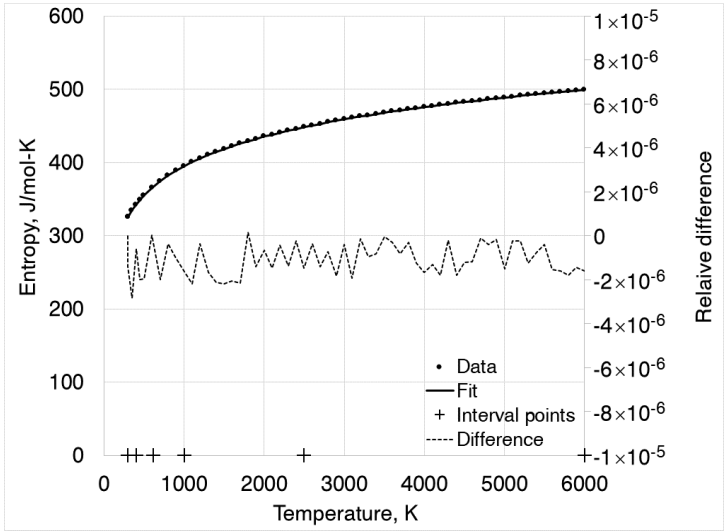


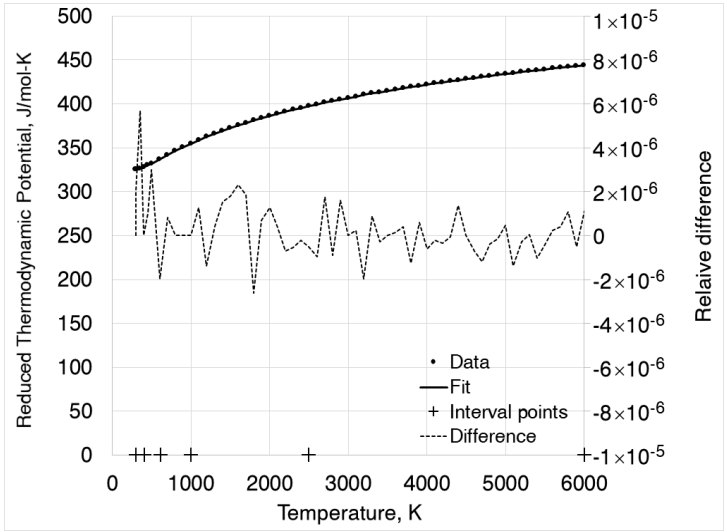
Barium Chloride ($BaCl_2$) – Ba1Cl2(g)

MW = 208.24 g/mol – $\Delta_f H^o$(298.15 K) = -498.733 kJ/mol – S^o(298.15 K) = 325.736 J/mol-K – p^o = 0.1 MPa



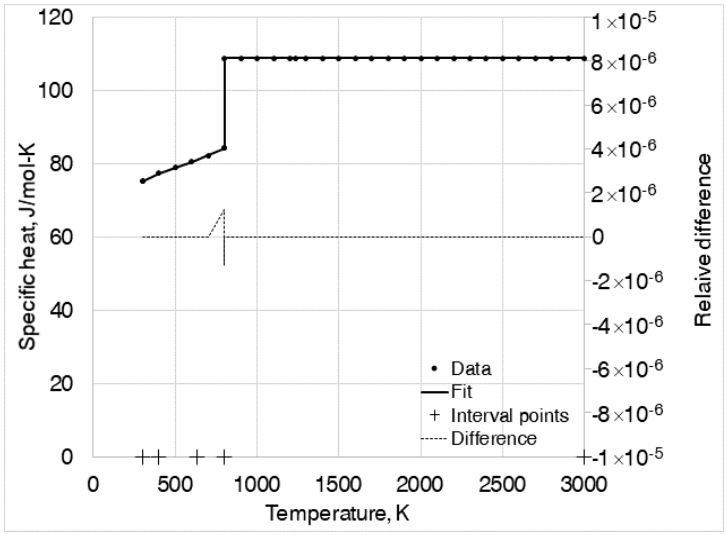


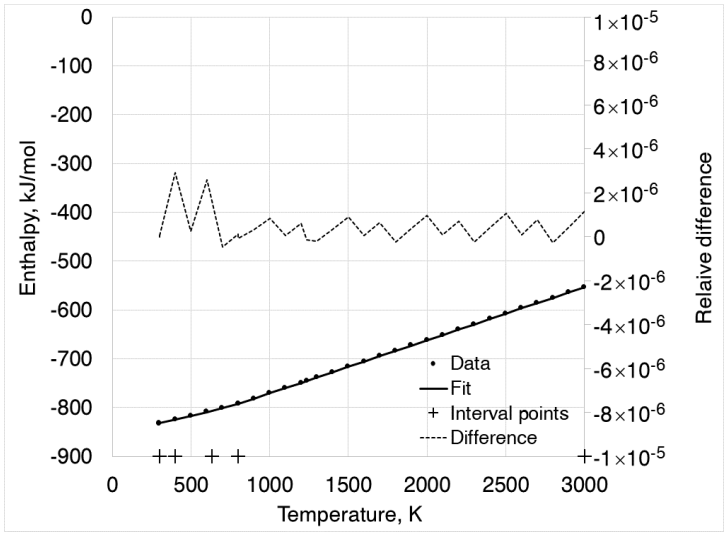


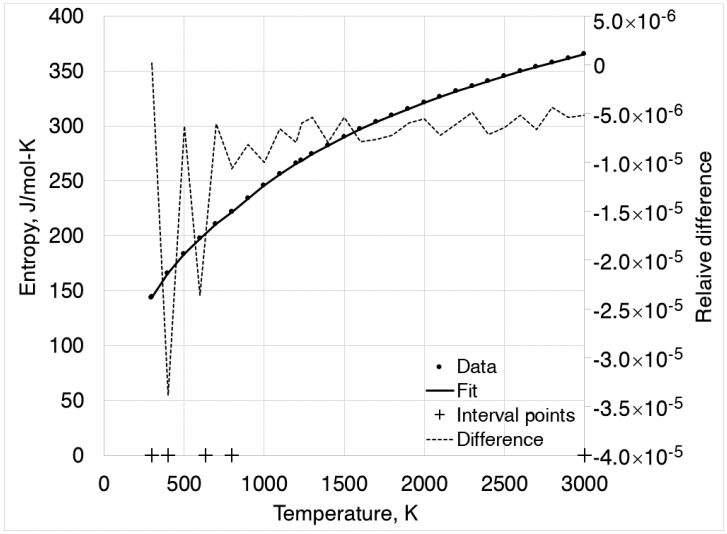


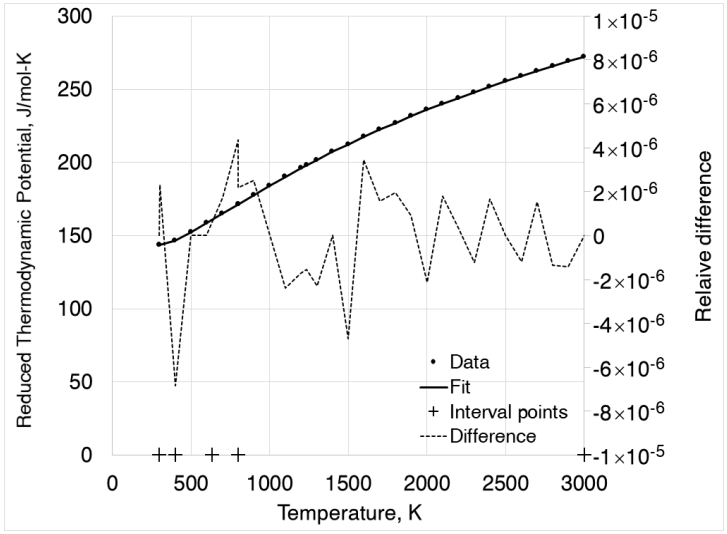
Barium Chloride ($BaCl_2$) – Ba1Cl2(l)

MW = 208.24 g/mol – $\Delta_f H^o$(298.15 K) = -832.467 kJ/mol – S^o(298.15 K) = 143.518 J/mol-K – p^o = 0.1 MPa



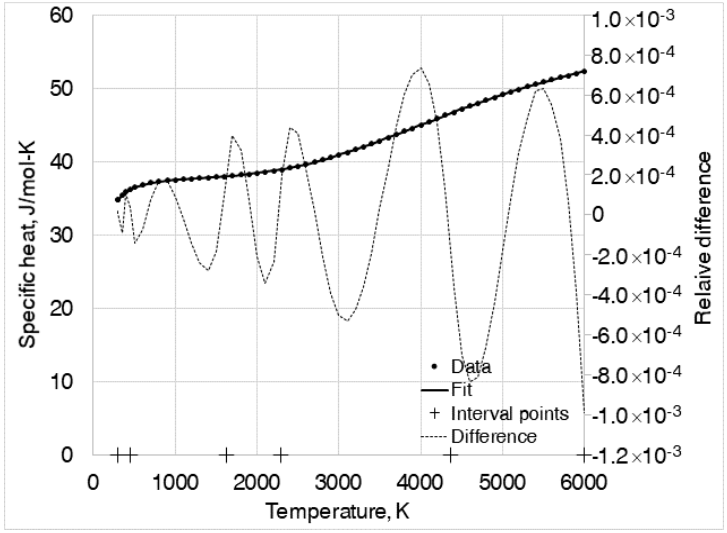


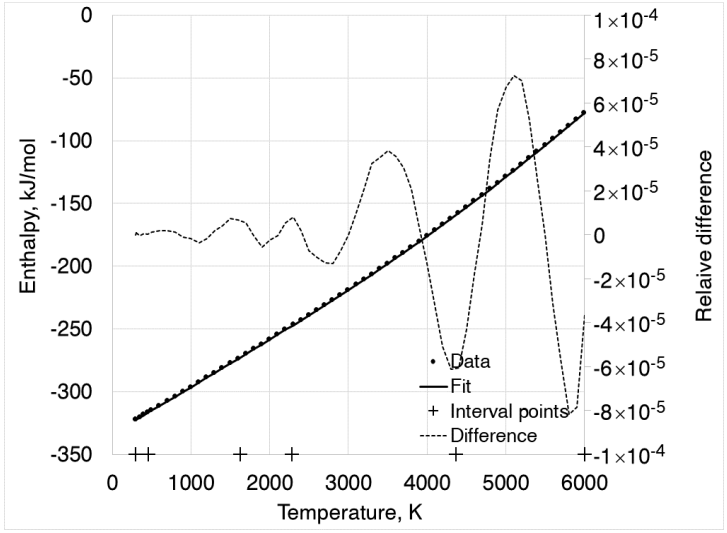


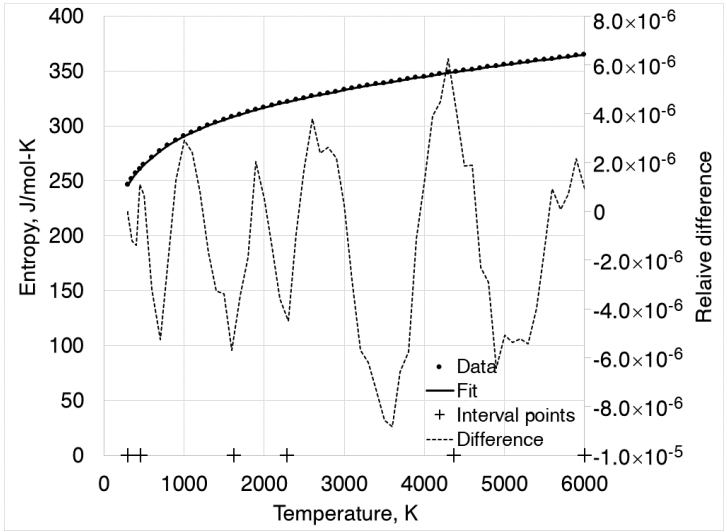


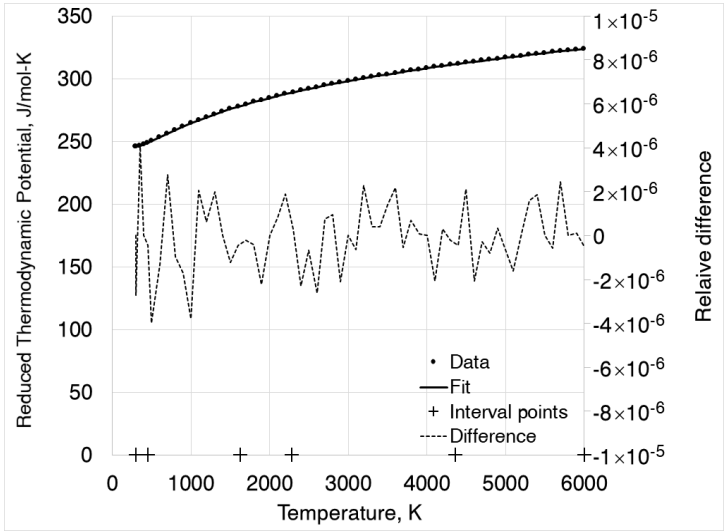
Barium Fluoride (BaF) – Ba1F1(g)

MW = 156.33 g/mol – $\Delta_f H^o$(298.15 K) = -322.168 kJ/mol – S^o(298.15 K) = 246.228 J/mol-K – p^o = 0.1 MPa



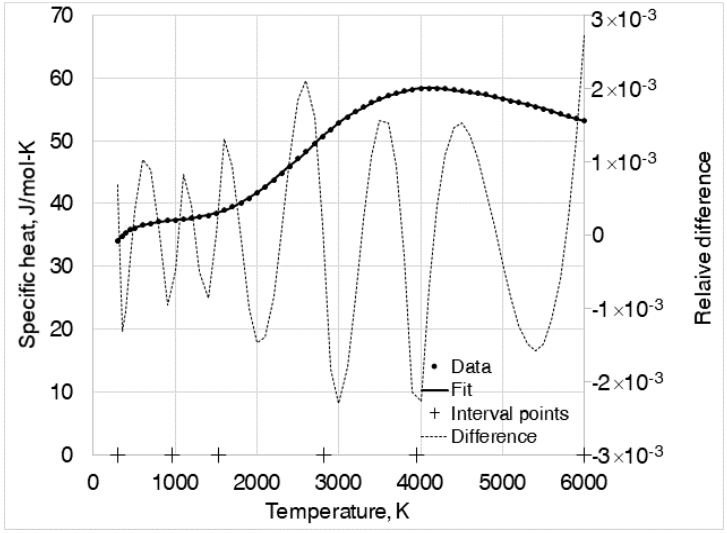


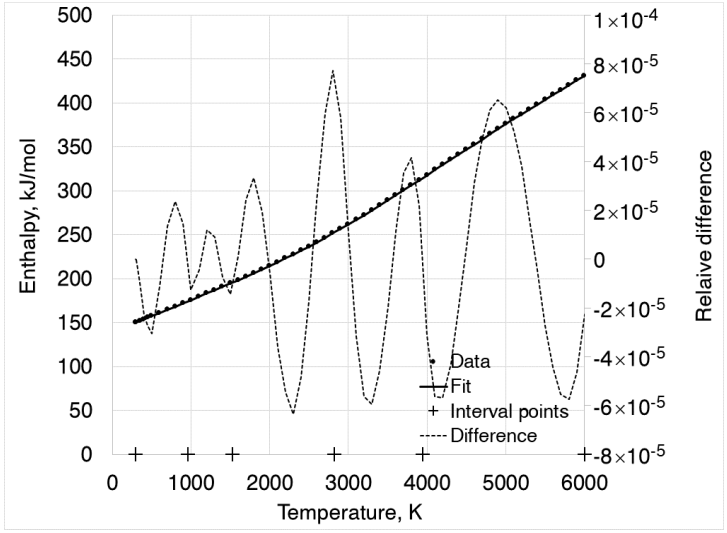


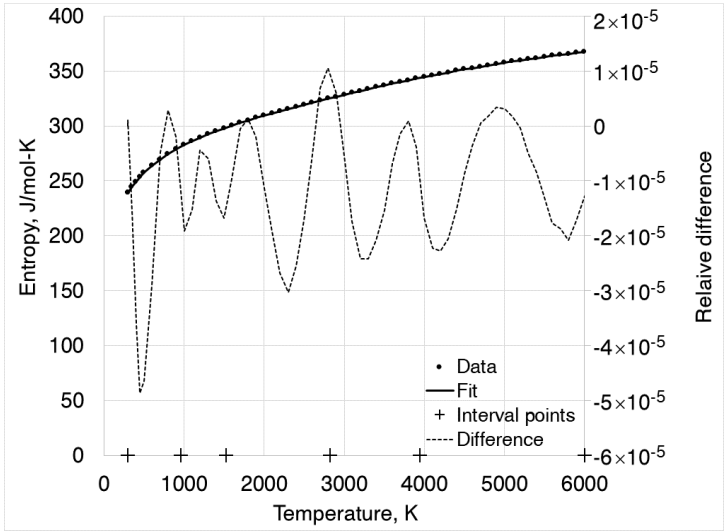


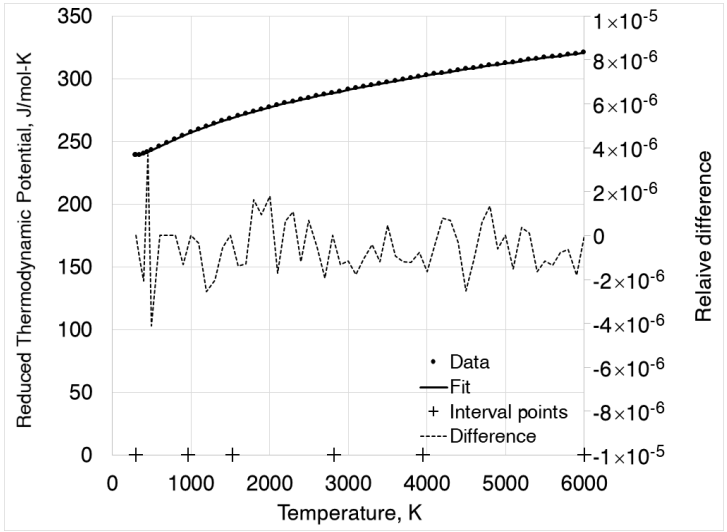
Barium Fluoride, Ion (BaF^+) – Ba1F1+(g)

MW = 156.33 g/mol – $\Delta_f H^o$(298.15 K) = 150.206 kJ/mol – S^o(298.15 K) = 238.955 J/mol-K – p^o = 0.1 MPa



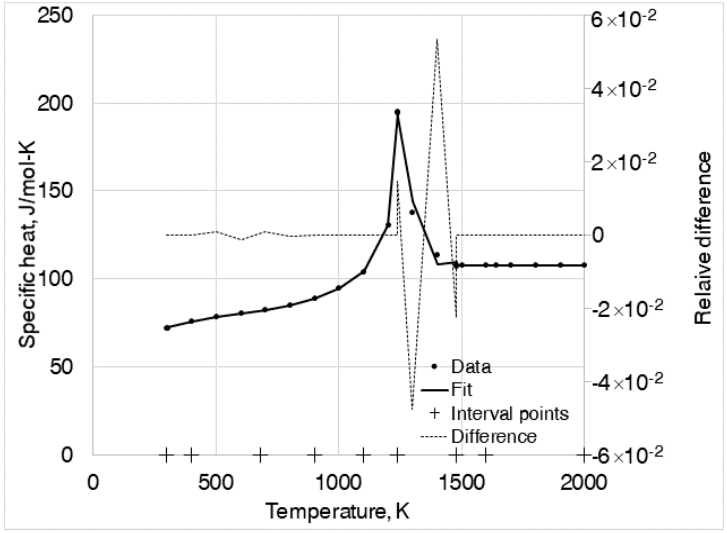


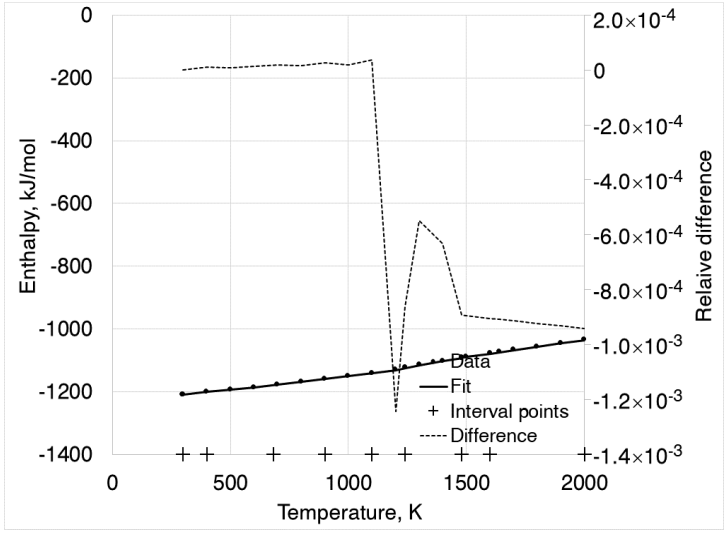


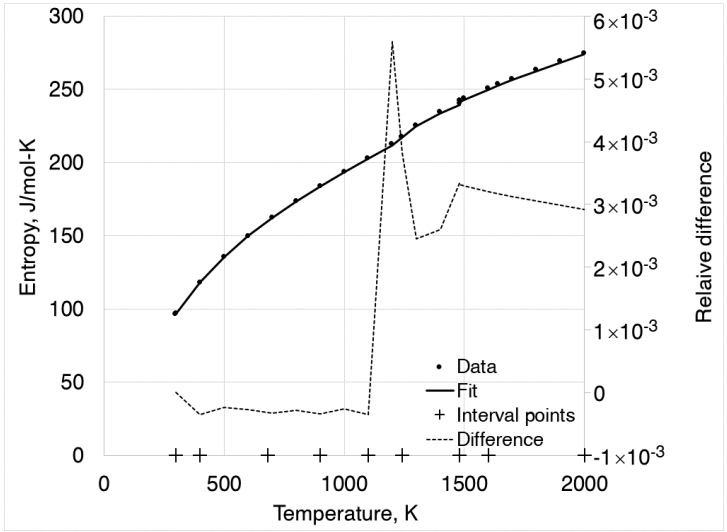


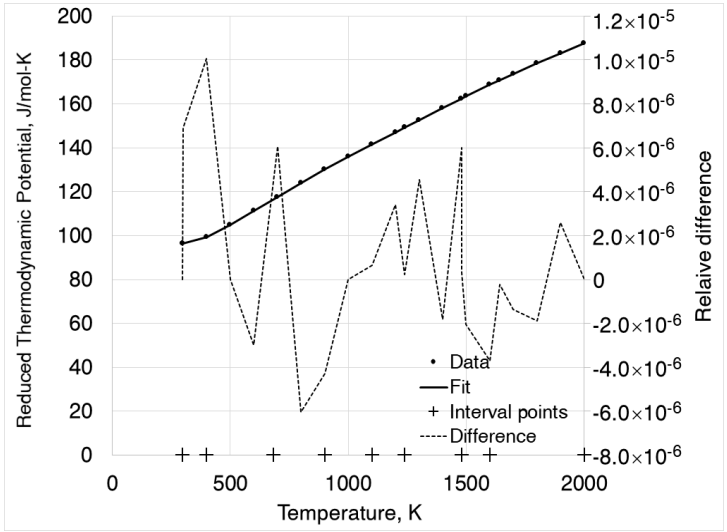
Barium Fluoride (BaF_2) – Ba1F2(cr)

MW = 175.33 g/mol – $\Delta_f H^o$(298.15 K) = -1208.758 kJ/mol – S^o(298.15 K) = 96.397 J/mol-K – p^o = 0.1 MPa



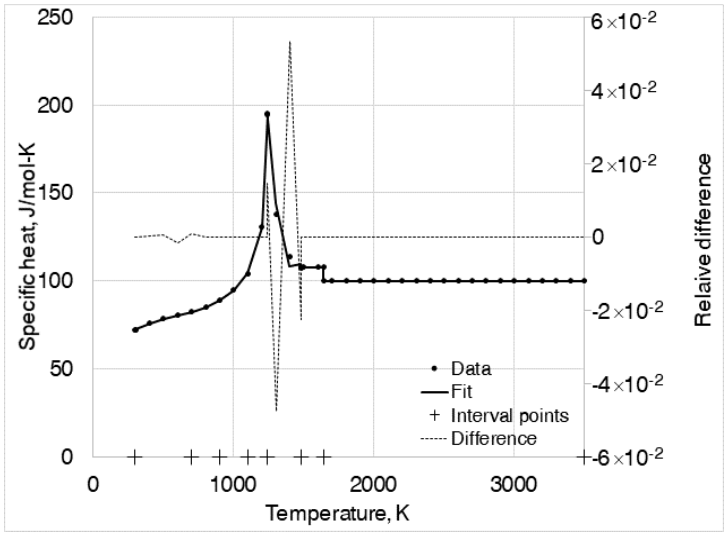


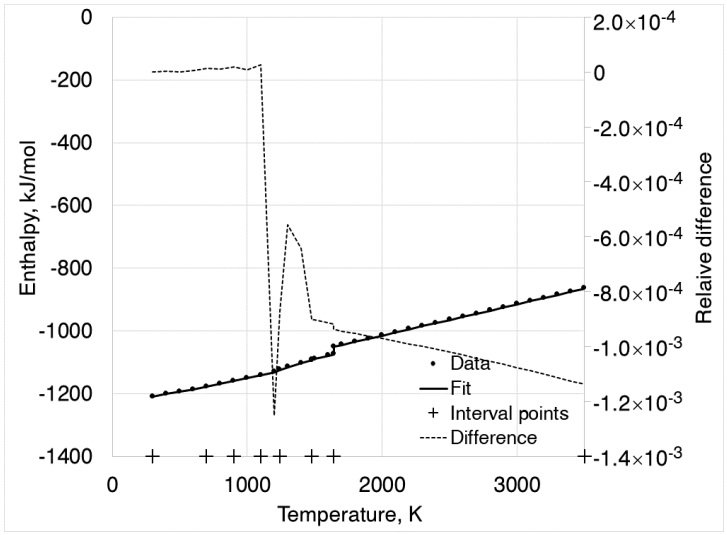


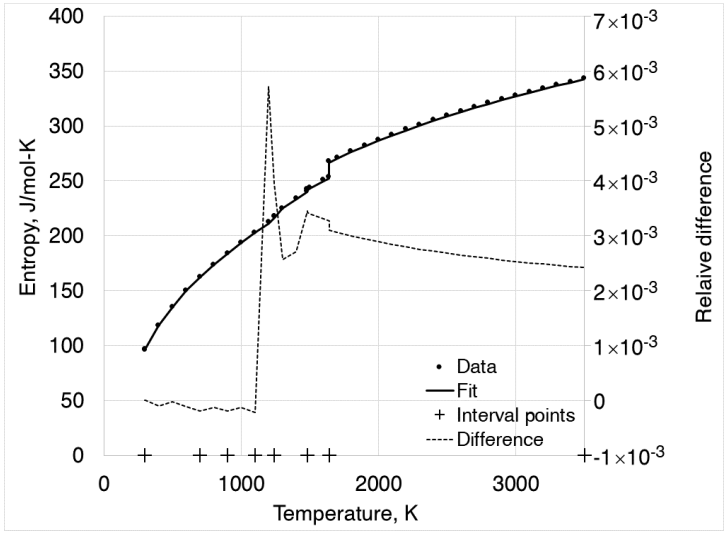


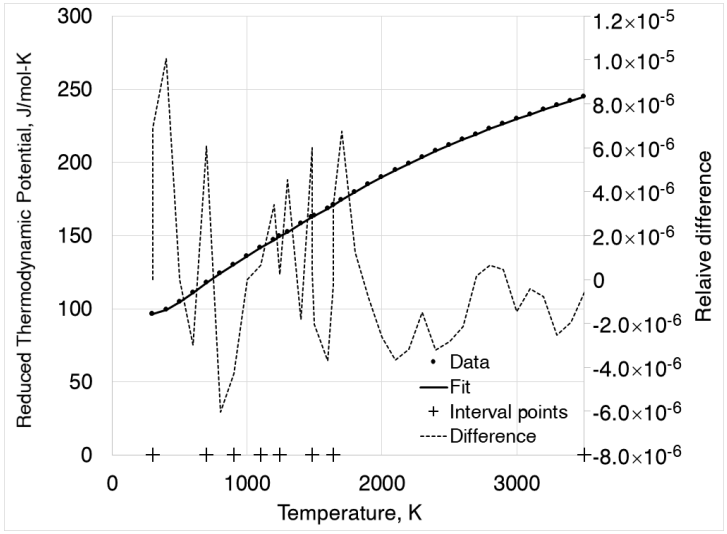
Barium Fluoride (BaF_2) – Ba1F2(cr,l)

MW = 175.33 g/mol – $\Delta_f H^o$(298.15 K) = -1208.758 kJ/mol – S^o(298.15 K) = 96.397 J/mol-K – p^o = 0.1 MPa



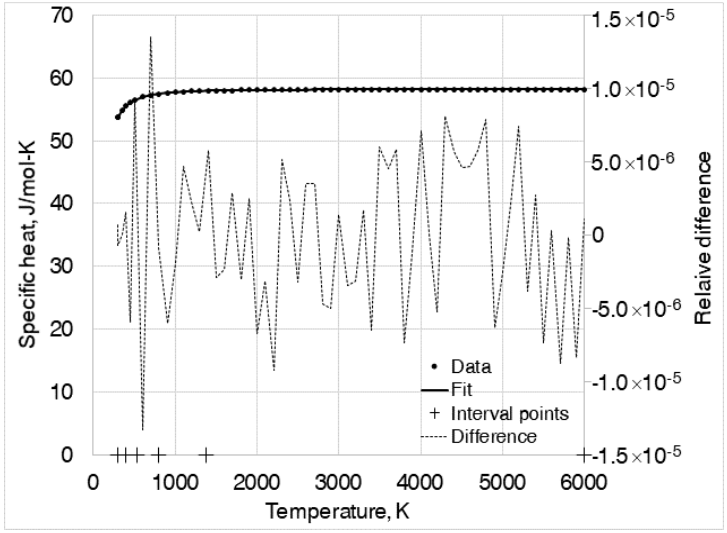


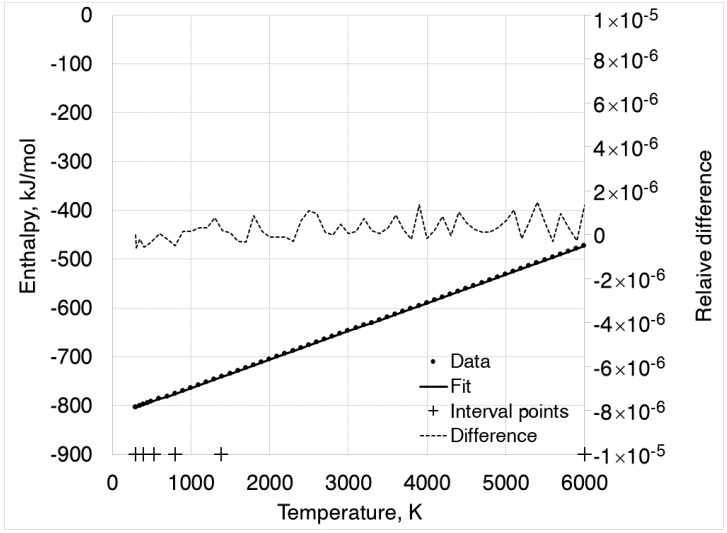


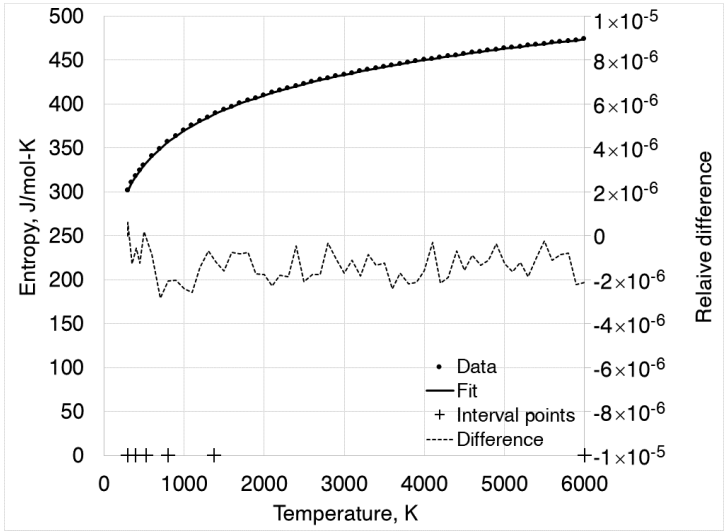


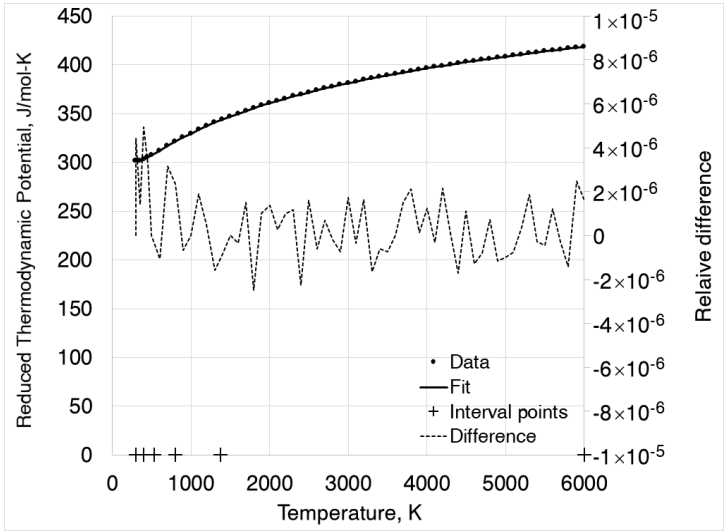
Barium Fluoride (BaF_2) – Ba1F2(g)

MW = 175.33 g/mol – $\Delta_f H^o$(298.15 K) = -803.746 kJ/mol – S^o(298.15 K) = 301.297 J/mol-K – p^o = 0.1 MPa



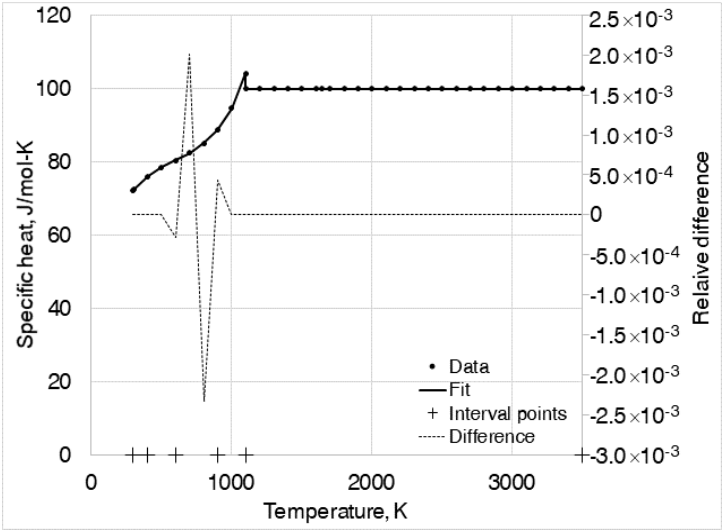


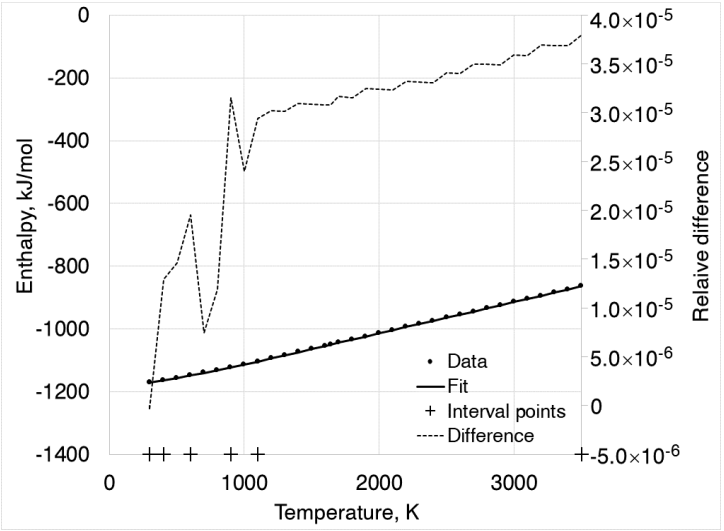


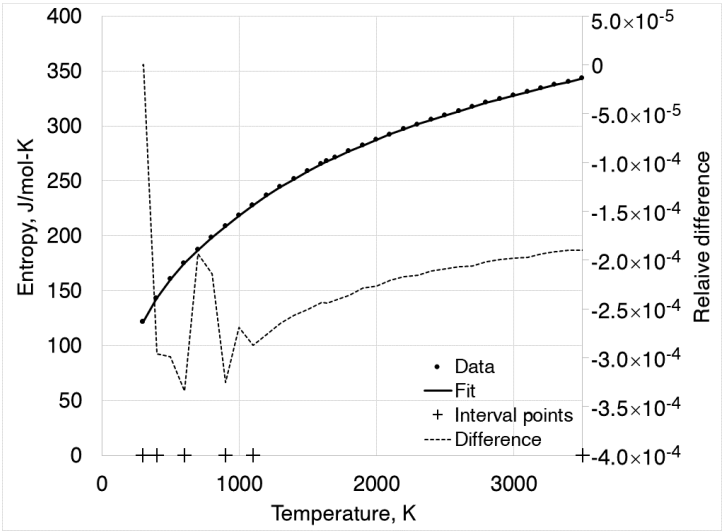


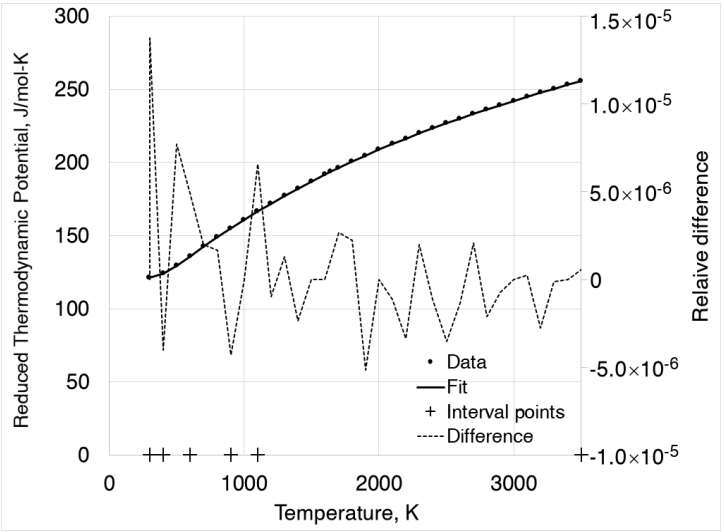
Barium Fluoride (BaF_2) – Ba1F2(l)

MW = 175.33 g/mol – $\Delta_f H^o$(298.15 K) = -1171.294 kJ/mol – S^o(298.15 K) = 121.254 J/mol-K – p^o = 0.1 MPa



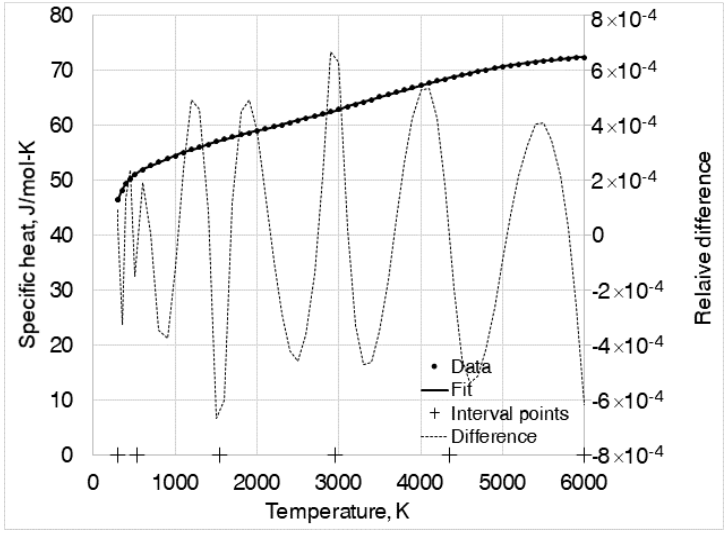


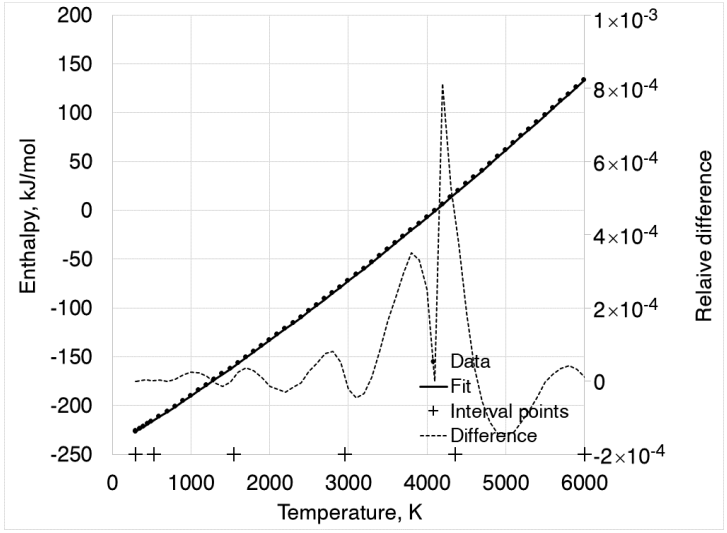


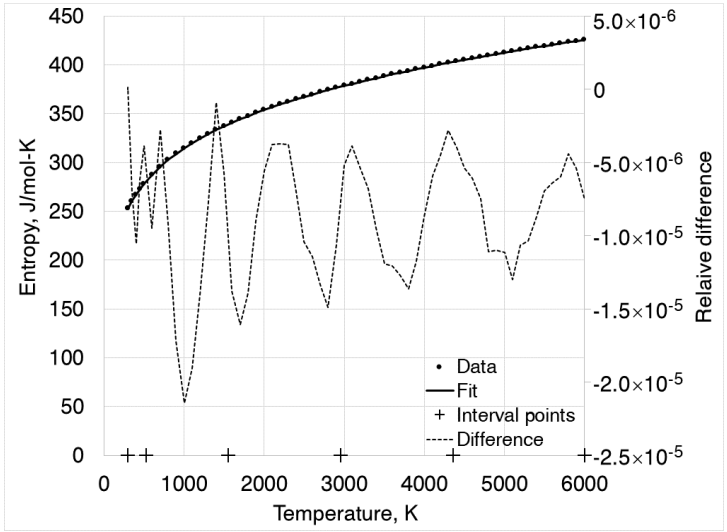


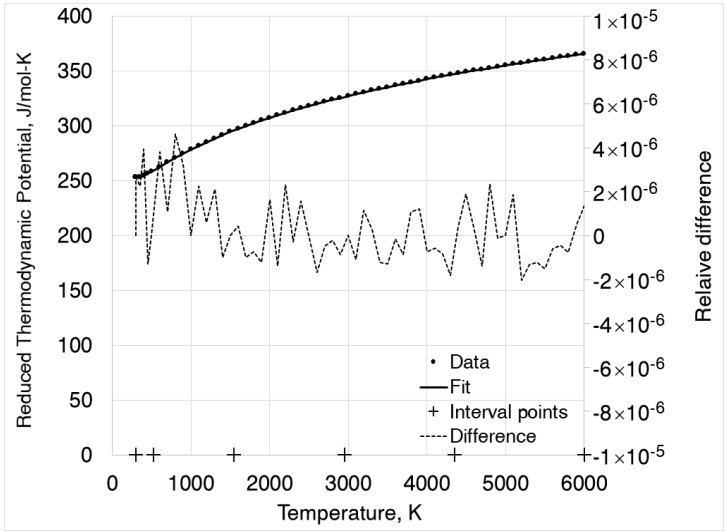
Barium Hydroxide (BaOH) – Ba1H1O1(g)

MW = 154.34 g/mol – $\Delta_f H^o$(298.15 K) = -226.44 kJ/mol – S^o(298.15 K) = 252.893 J/mol-K – p^o = 0.1 MPa



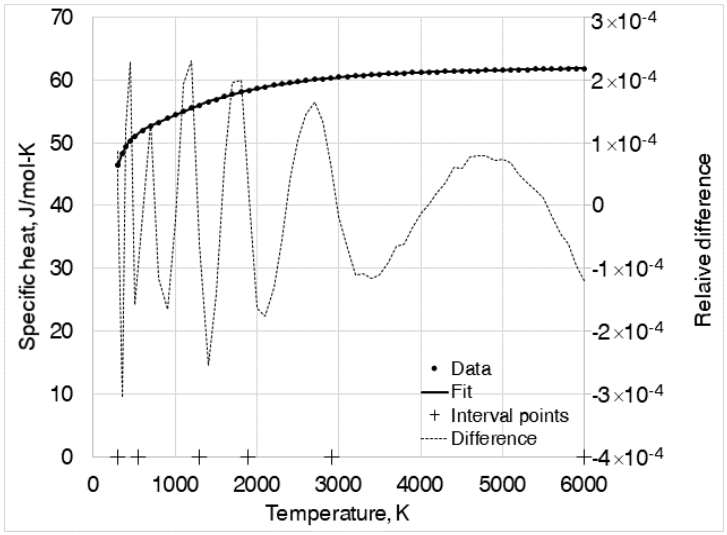


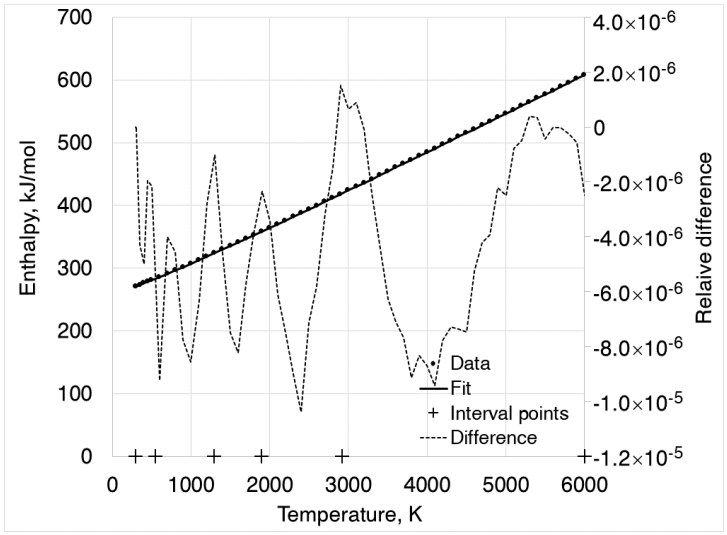


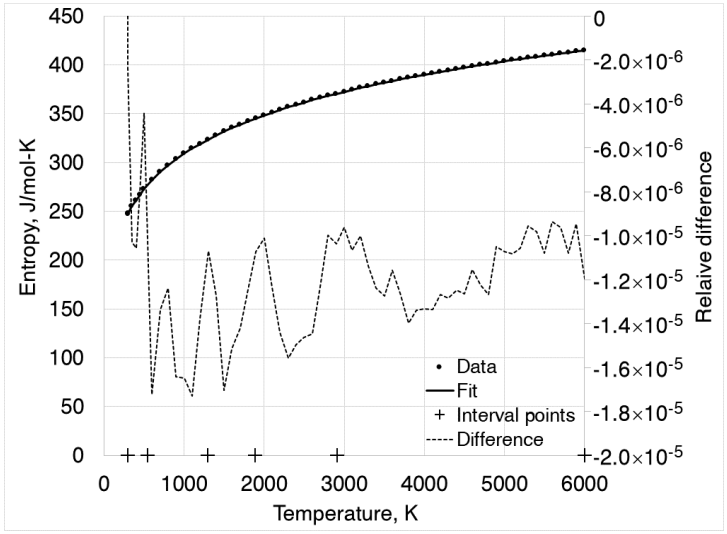


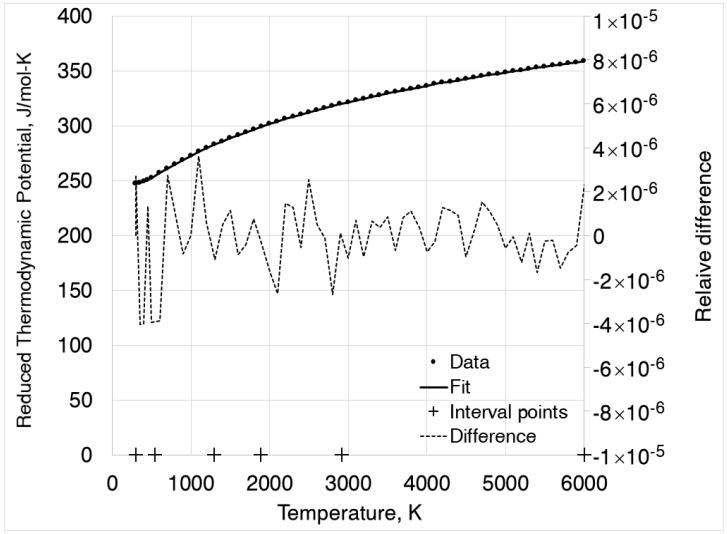
Barium Hydroxide, Ion ($BaOH^+$) – Ba1H1O1+(g)

MW = 154.34 g/mol – $\Delta_f H^o$(298.15 K) = 270.897 kJ/mol – S^o(298.15 K) = 247.265 J/mol-K – p^o = 0.1 MPa



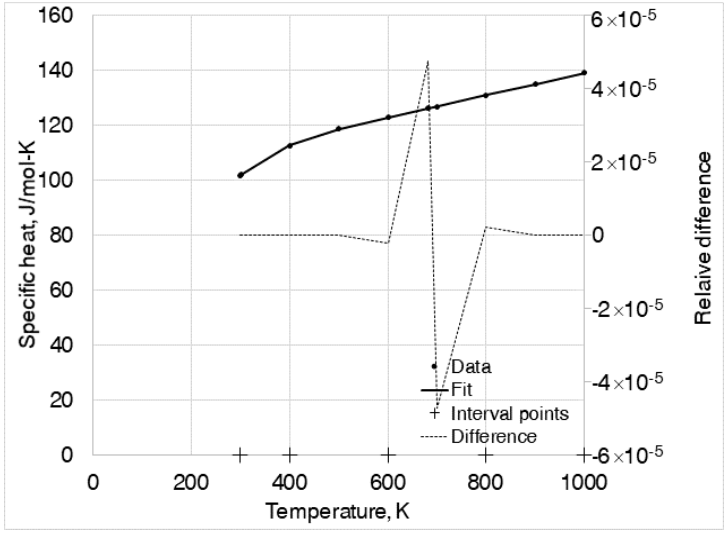


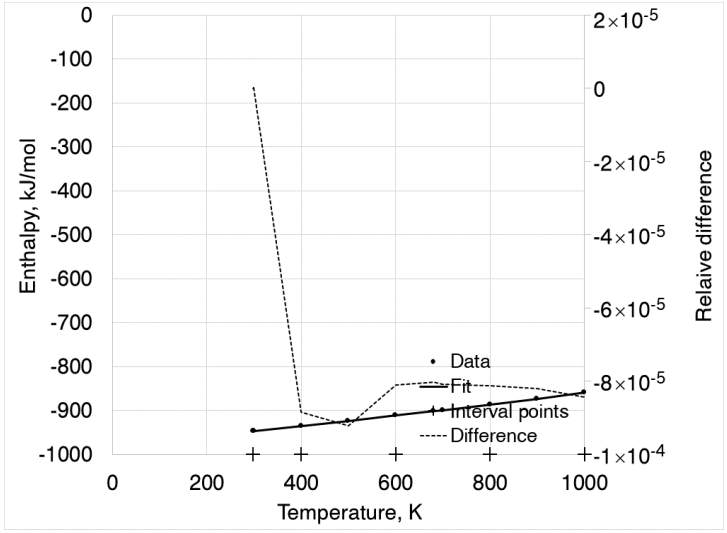


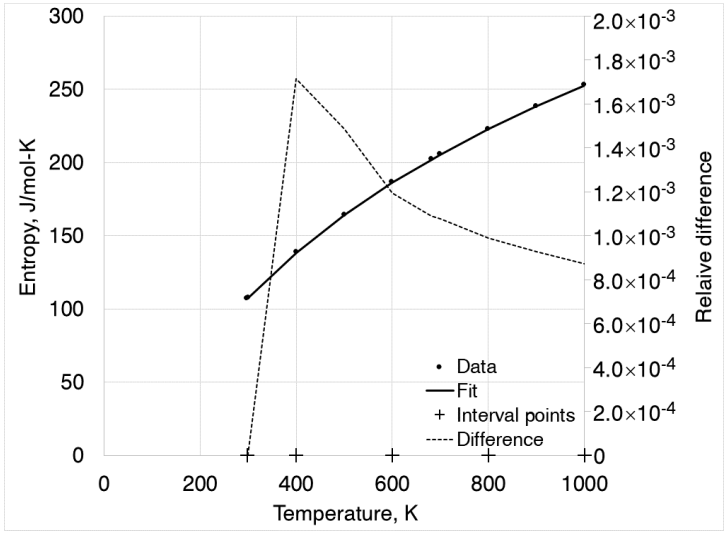


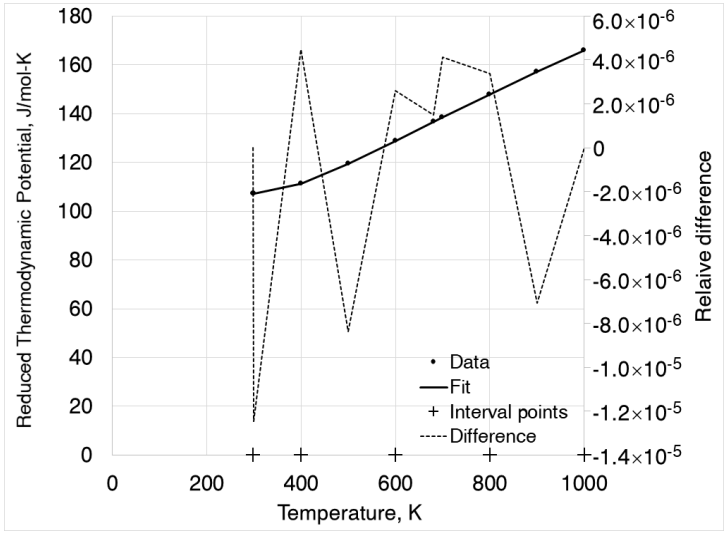
Barium Hydroxide, Alpha ($Ba(OH)_2$) – Ba1H2O2(cr)

MW = 171.34 g/mol – $\Delta_f H^o$(298.15 K) = -946.295 kJ/mol – S^o(298.15 K) = 107.11 J/mol-K – p^o = 0.1 MPa



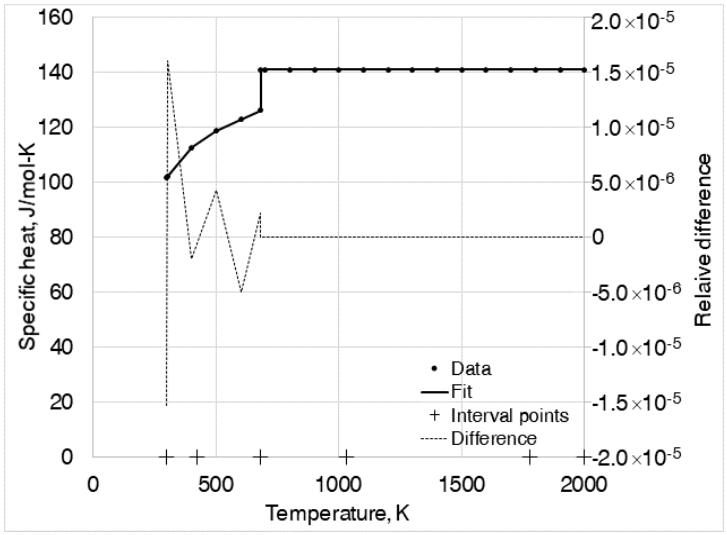


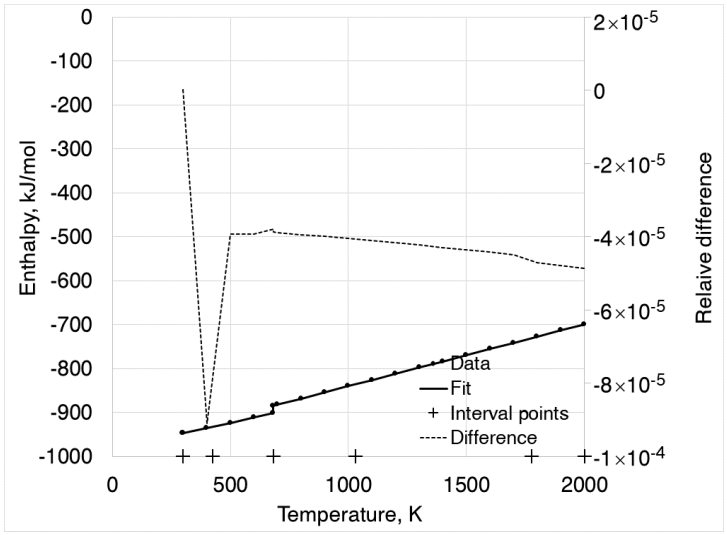


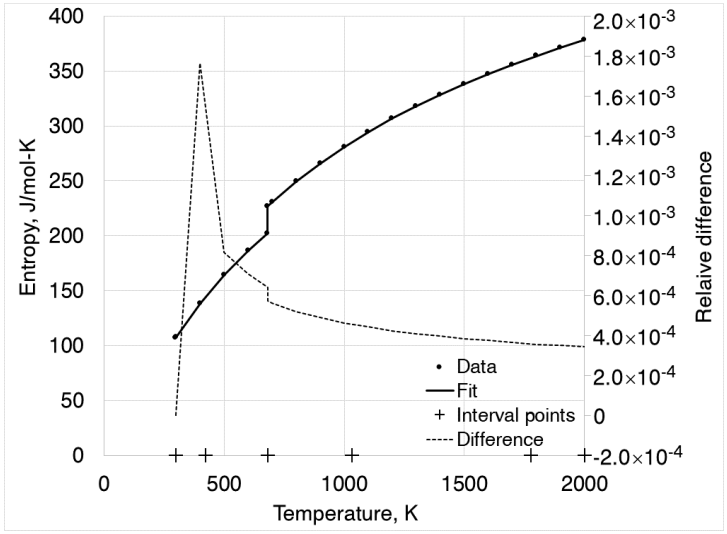


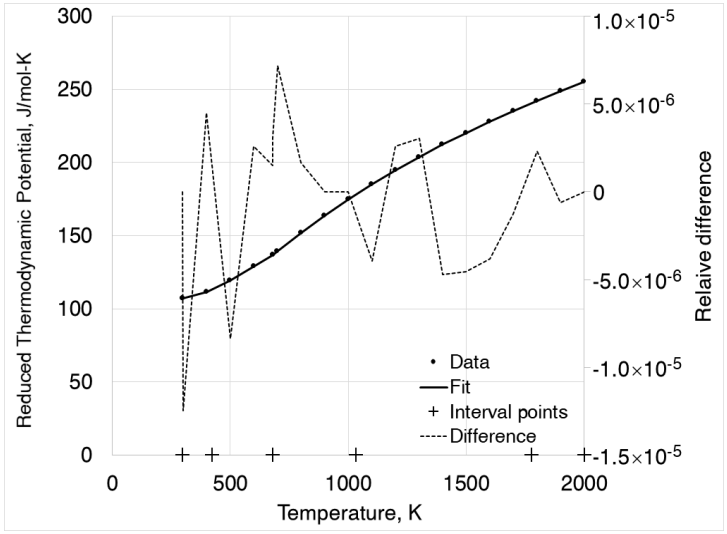
Barium Hydroxide ($Ba(OH)_2$) – Ba1H2O2(cr,l)

MW = 171.34 g/mol – $\Delta_f H^o$(298.15 K) = -946.295 kJ/mol – S^o(298.15 K) = 107.11 J/mol-K – p^o = 0.1 MPa



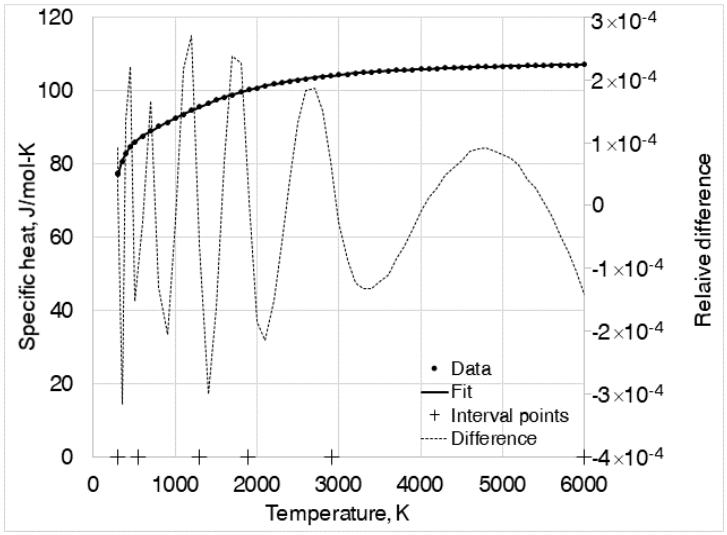


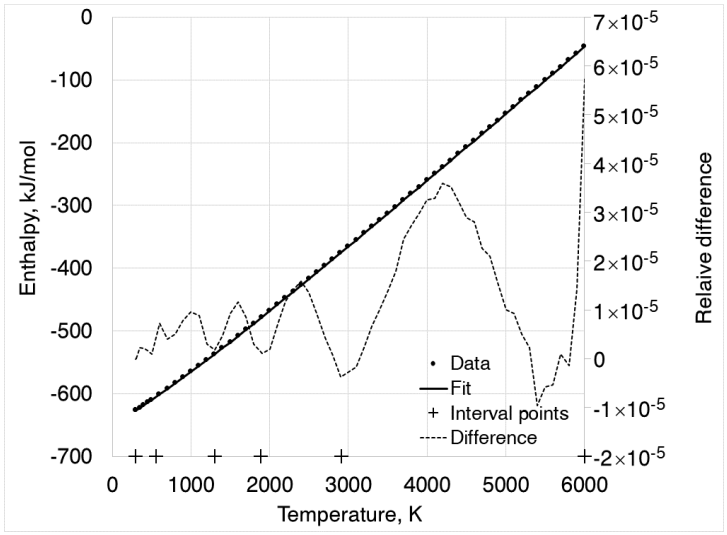


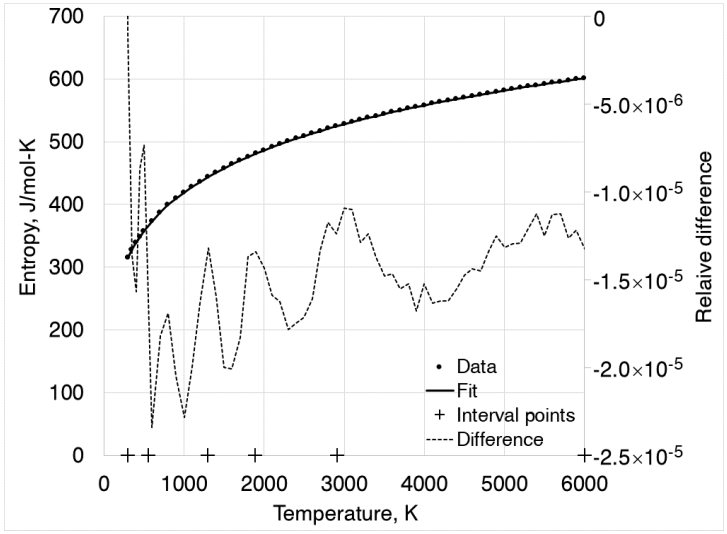


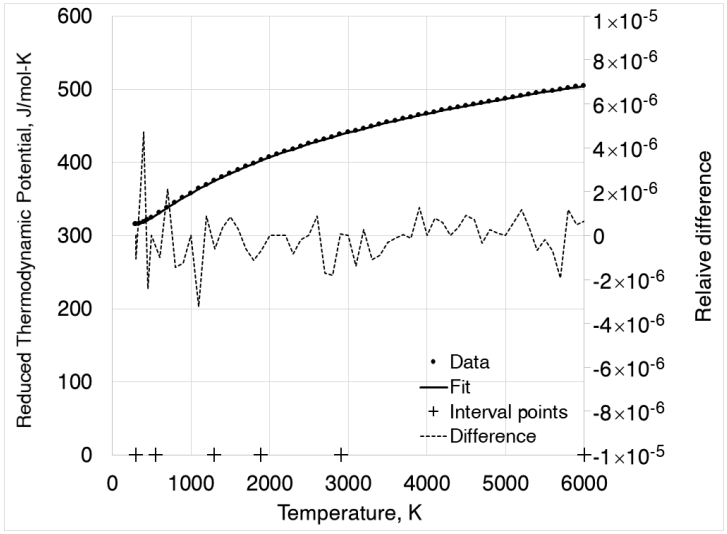
Barium Hydroxide ($Ba(OH)_2$) – Ba1H2O2(g)

MW = 171.34 g/mol – $\Delta_f H^o$(298.15 K) = -626.562 kJ/mol – S^o(298.15 K) = 315.016 J/mol-K – p^o = 0.1 MPa



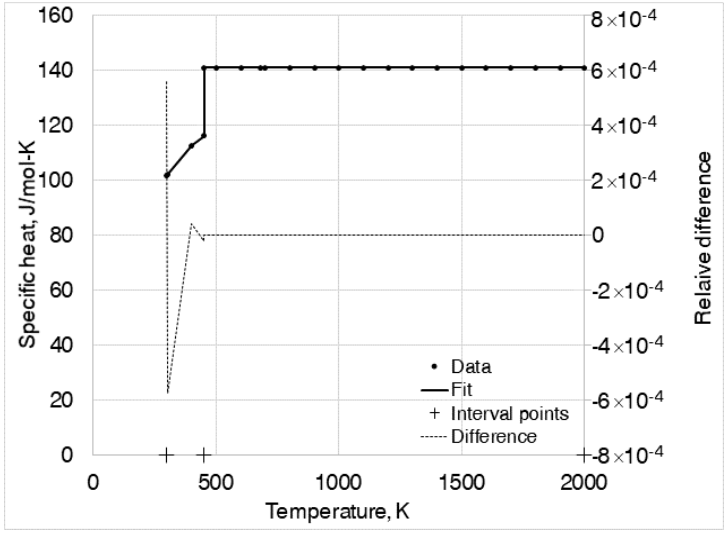


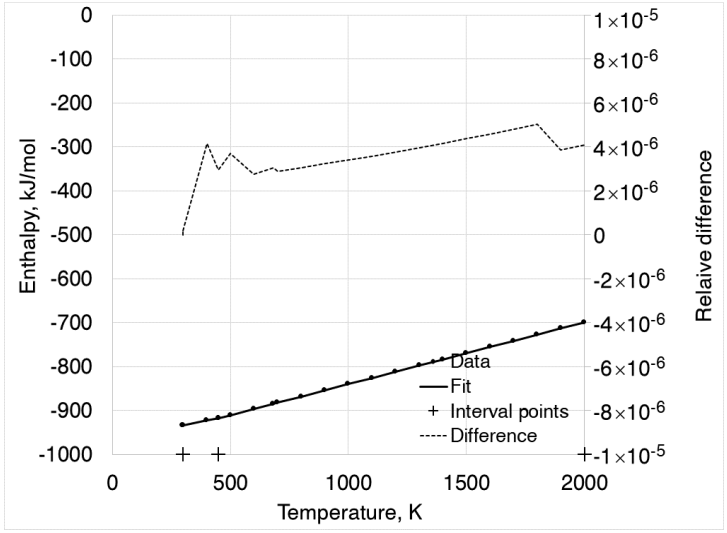


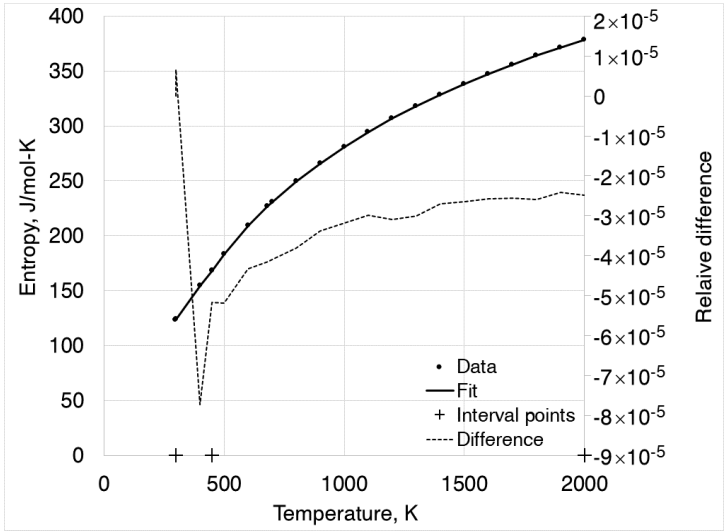


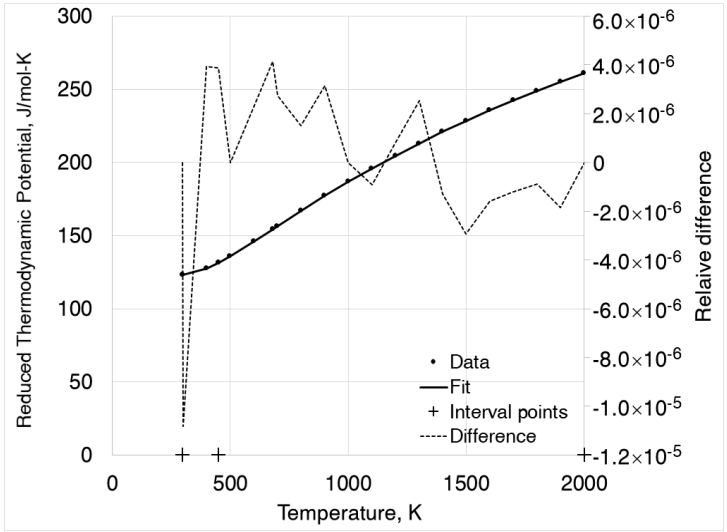
Barium Hydroxide ($Ba(OH)_2$) – Ba1H2O2(l)

MW = 171.34 g/mol – $\Delta_f H^o$(298.15 K) = -934.123 kJ/mol – S^o(298.15 K) = 123.357 J/mol-K – p^o = 0.1 MPa



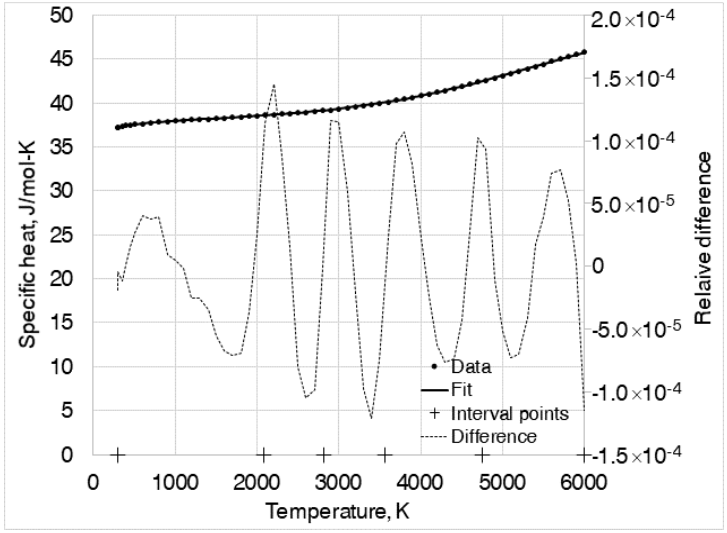


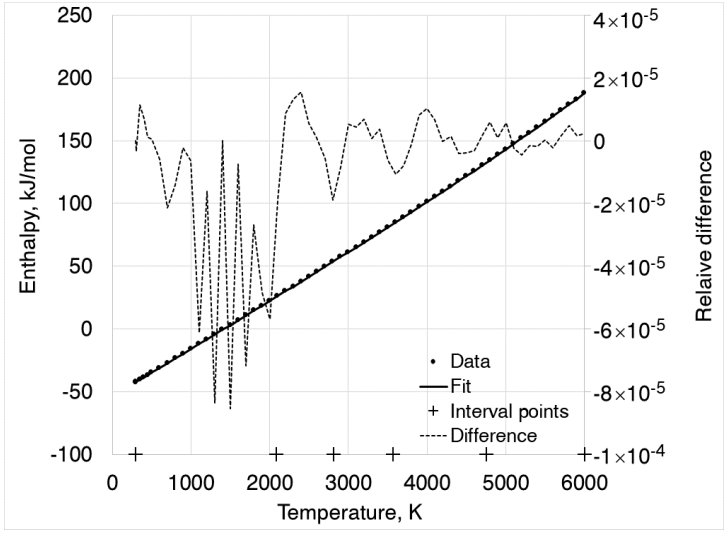


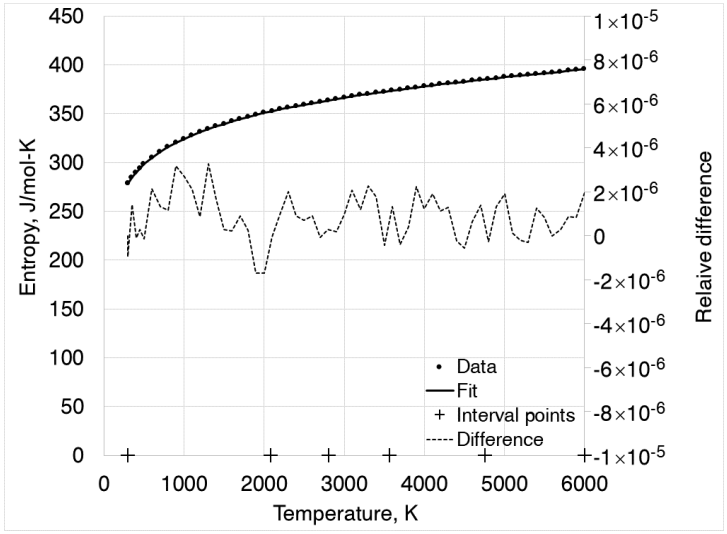


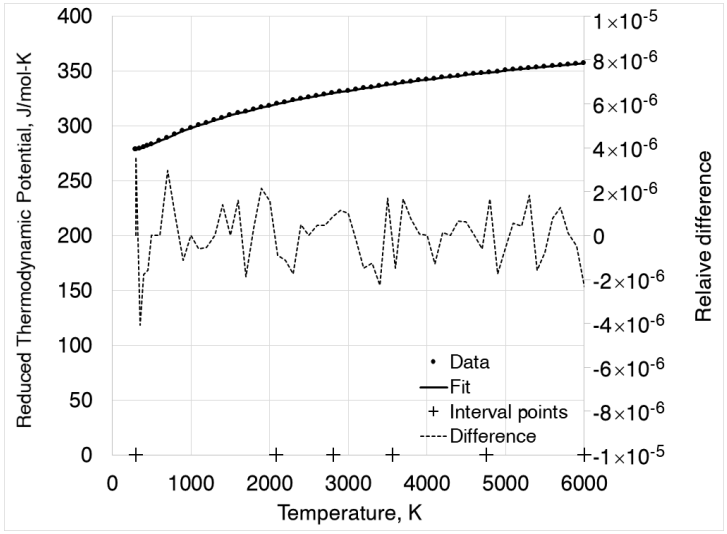
Barium Iodide (BaI) – Ba1I1(g)

MW = 264.23 g/mol – $\Delta_f H^o$(298.15 K) = -42.417 kJ/mol – S^o(298.15 K) = 278.737 J/mol-K – p^o = 0.1 MPa



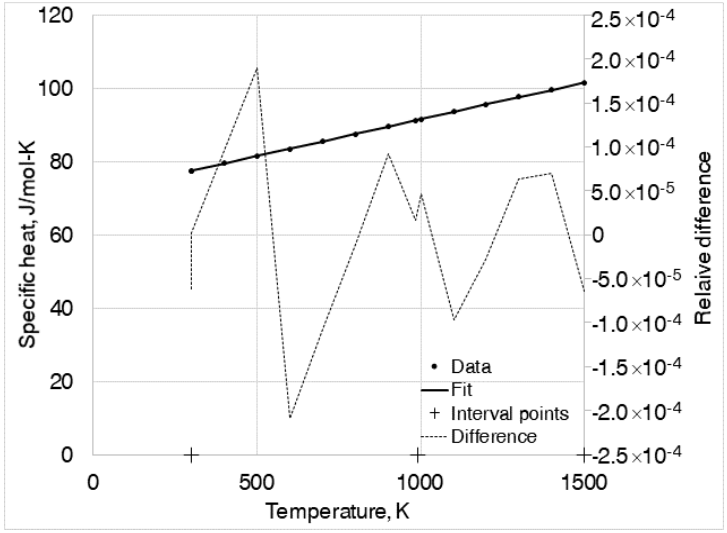


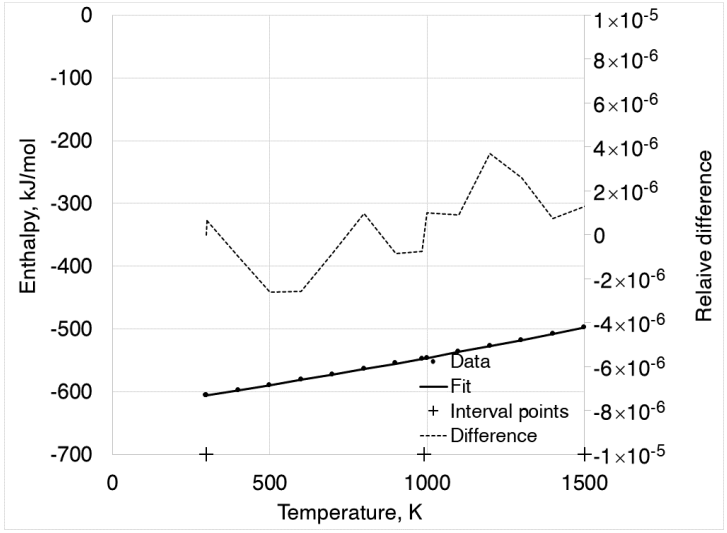


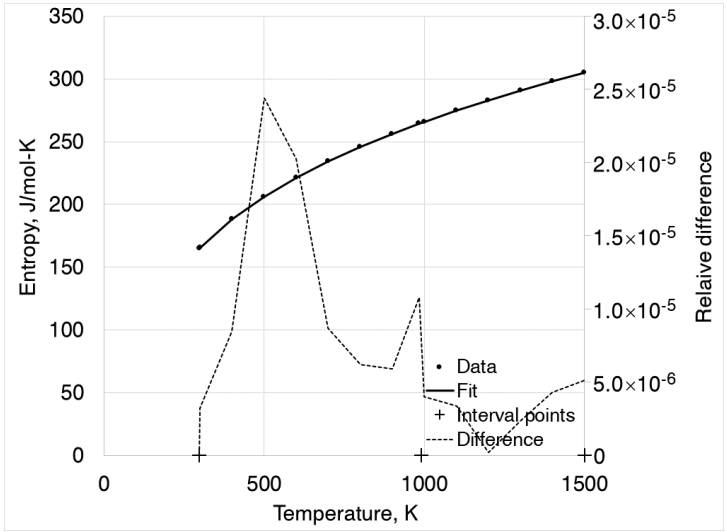


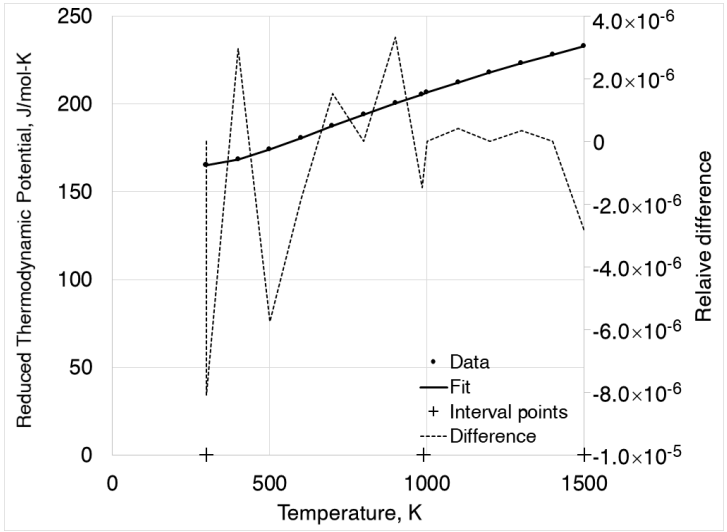
Barium Iodide (BaI_2) – Ba1I2(cr)

MW = 391.14 g/mol – $\Delta_f H^o$(298.15 K) = -605.425 kJ/mol – S^o(298.15 K) = 165.142 J/mol-K – p^o = 0.1 MPa



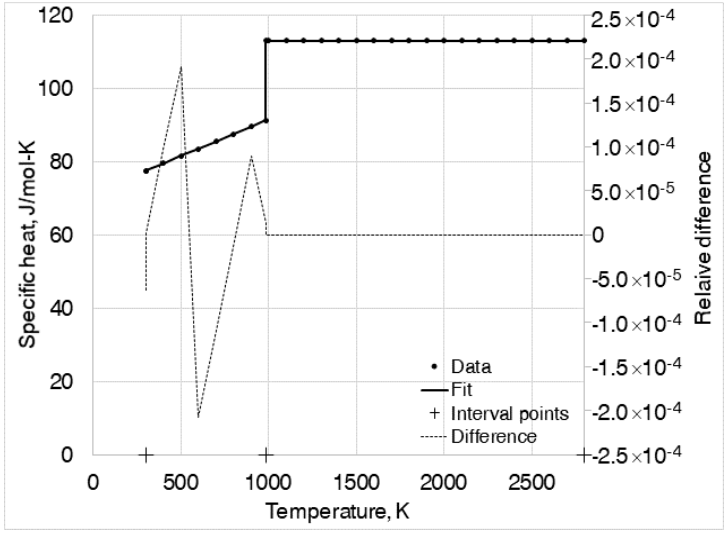


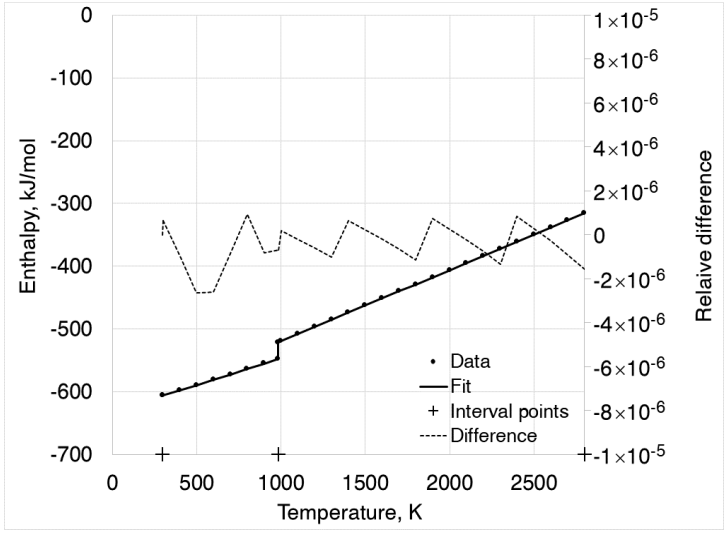


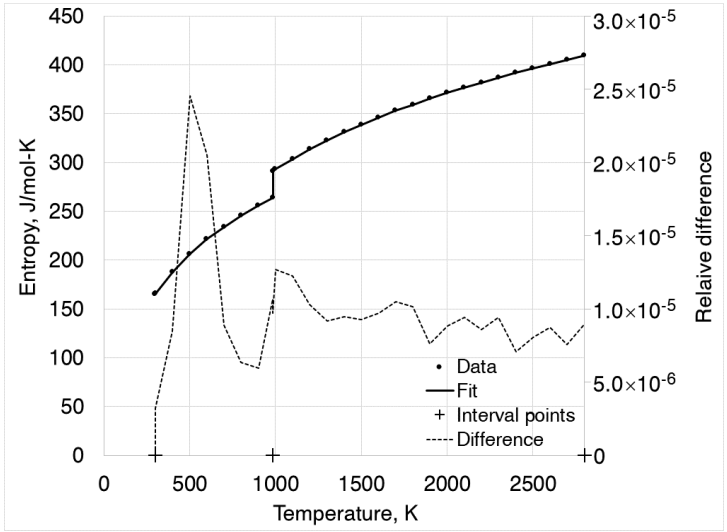


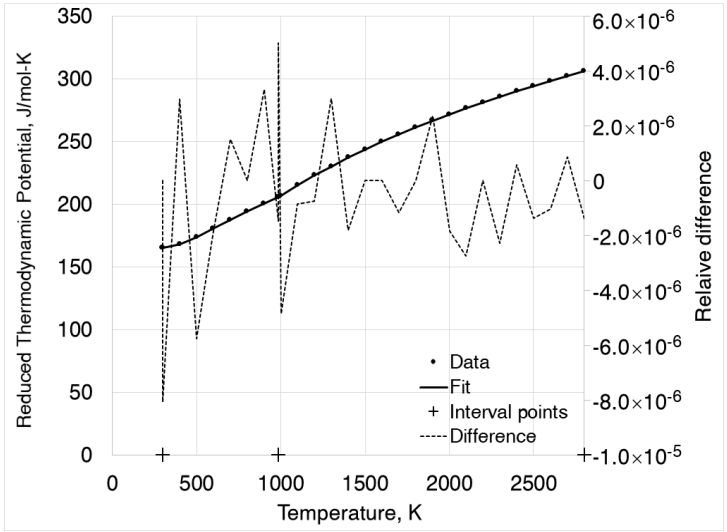
Barium Iodide (BaI_2) – Ba1I2(cr,l)

MW = 391.14 g/mol – $\Delta_f H^o$(298.15 K) = -605.425 kJ/mol – S^o(298.15 K) = 165.142 J/mol-K – p^o = 0.1 MPa



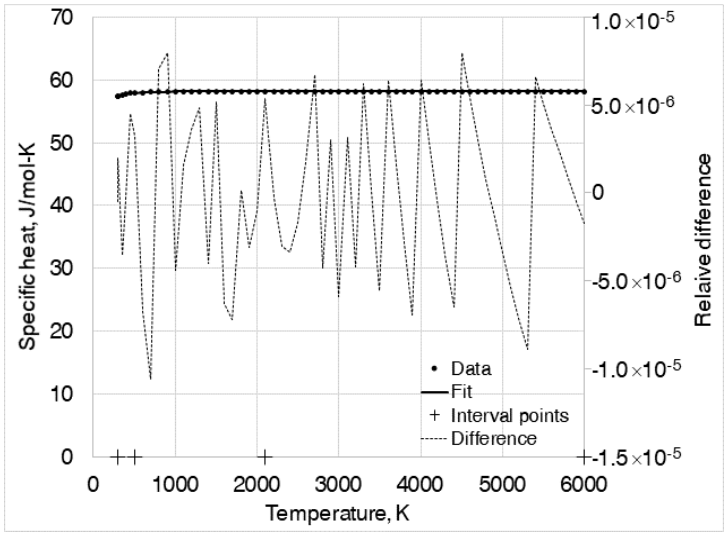


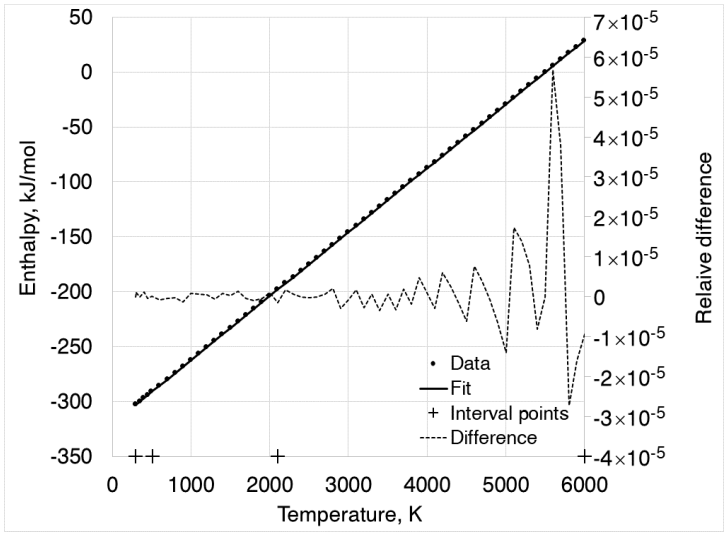


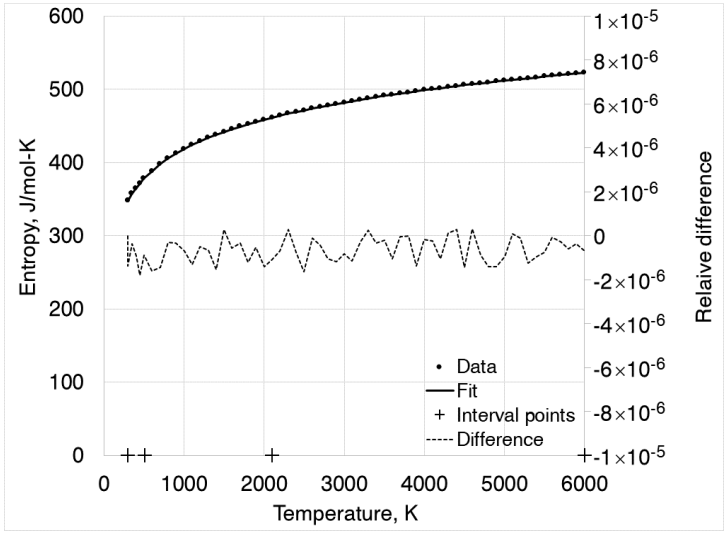


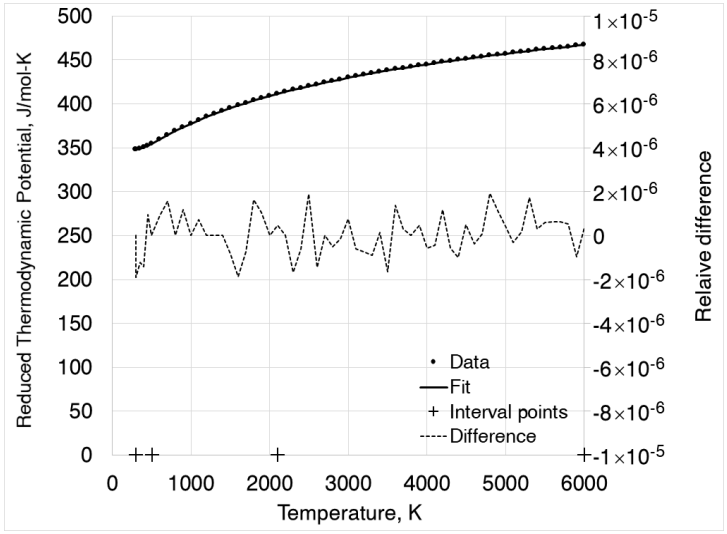
Barium Iodide (BaI_2) – Ba1I2(g)

MW = 391.14 g/mol – $\Delta_f H^o$(298.15 K) = -302.922 kJ/mol – S^o(298.15 K) = 348.139 J/mol-K – p^o = 0.1 MPa



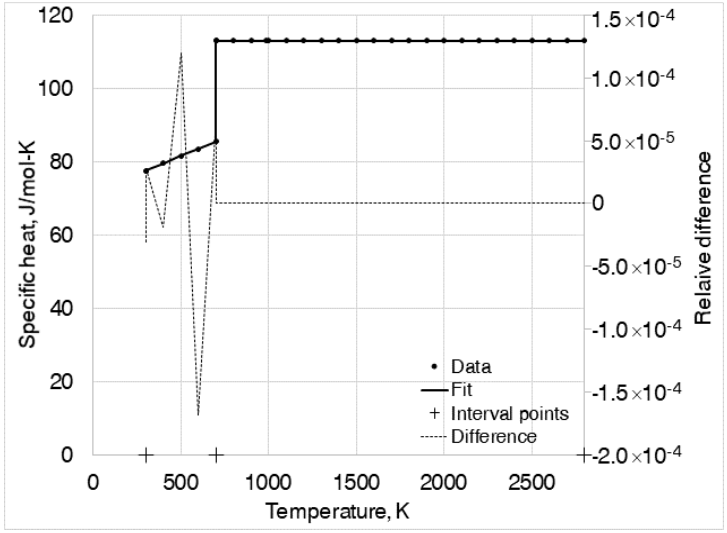


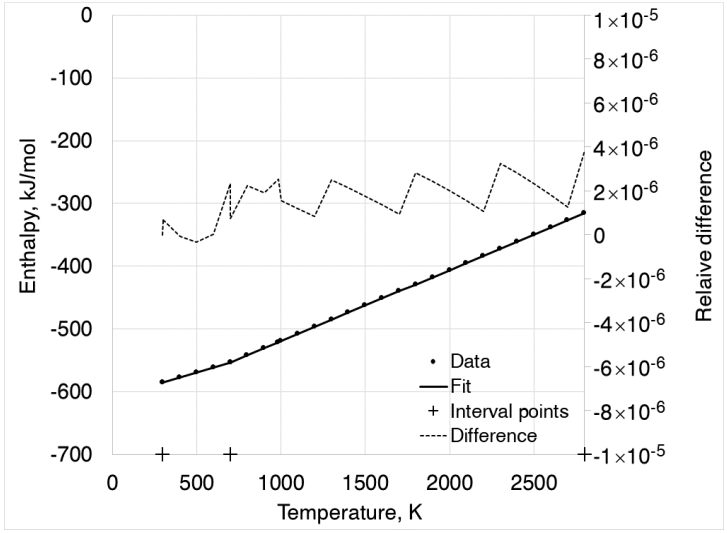


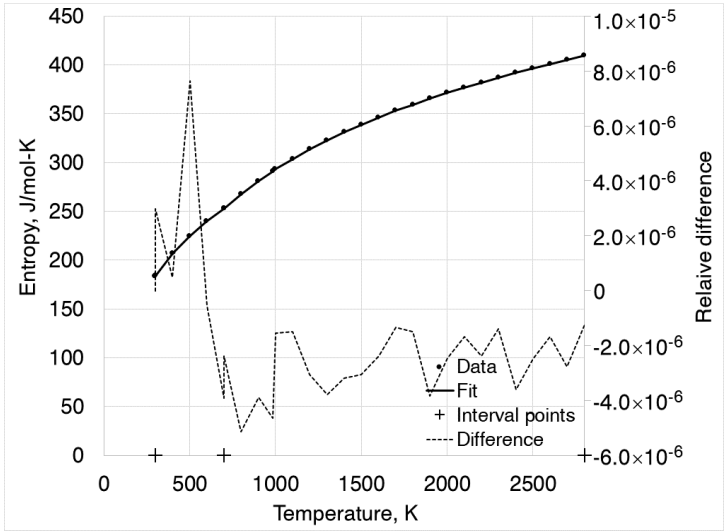


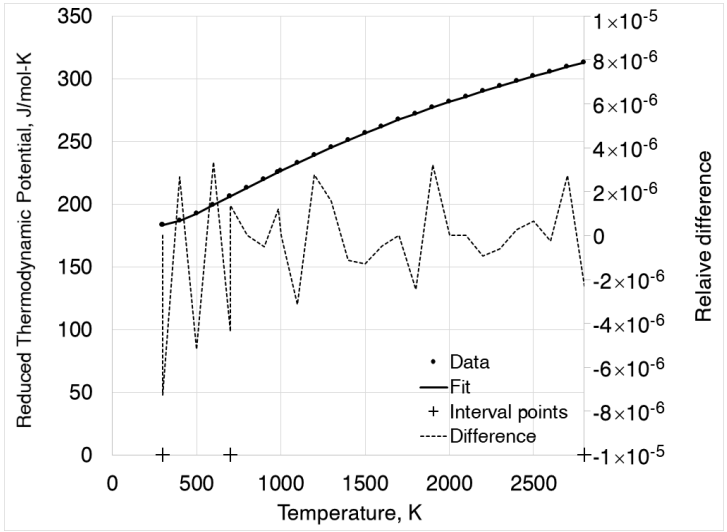
Barium Iodide (BaI_2) – Ba1I2(l)

MW = 391.14 g/mol – $\Delta_f H^o$(298.15 K) = -585.884 kJ/mol – S^o(298.15 K) = 183.669 J/mol-K – p^o = 0.1 MPa



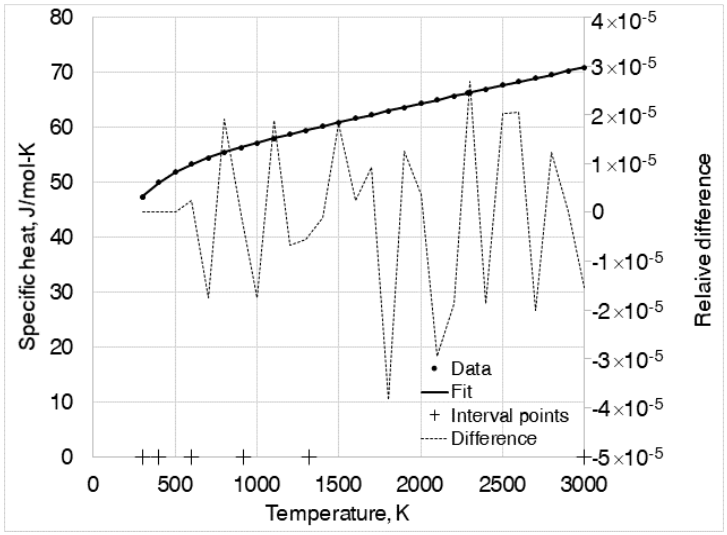


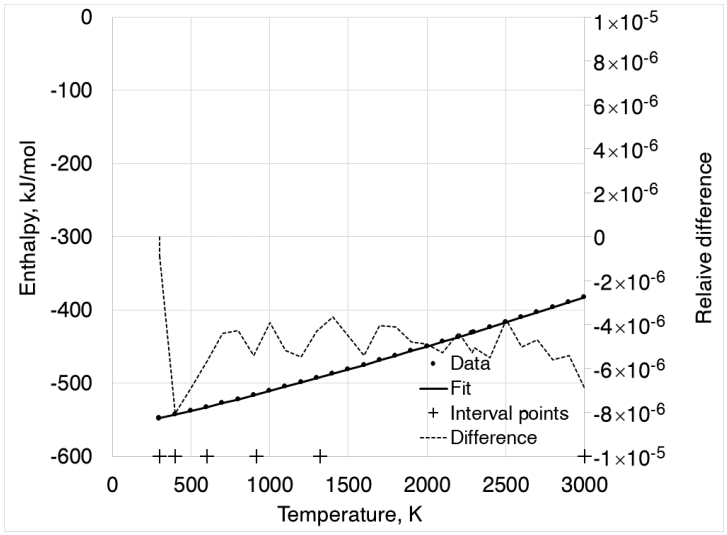


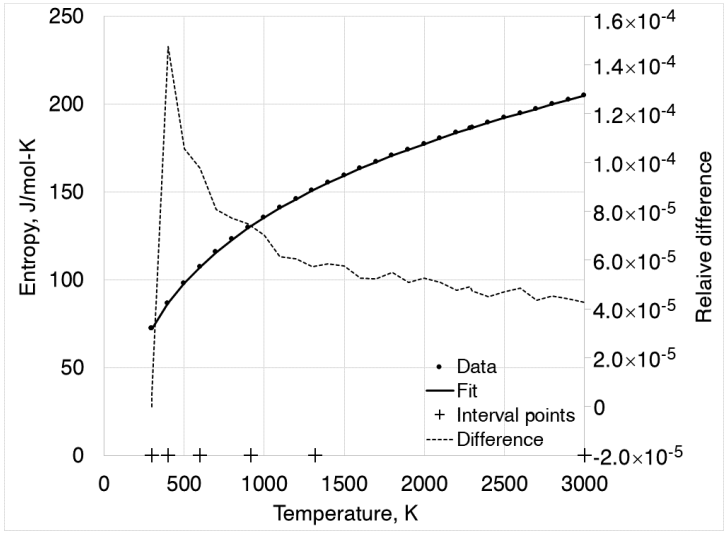


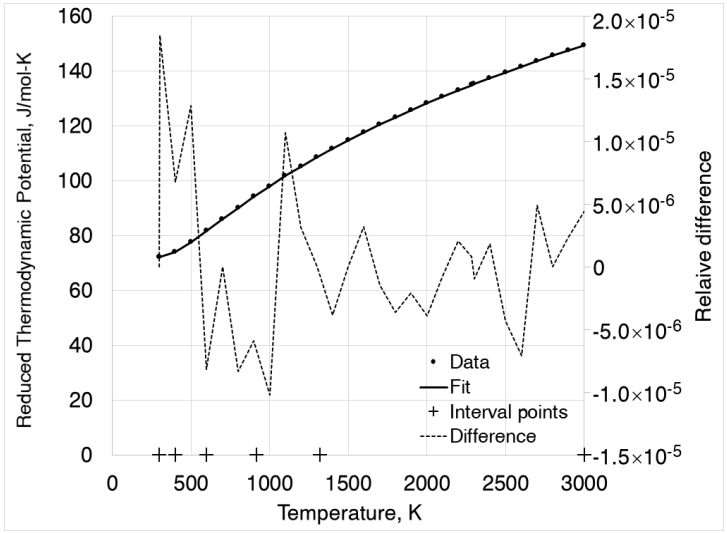
Barium Oxide (BaO) – Ba1O1(cr)

MW = 153.33 g/mol – $\Delta_f H^o$(298.15 K) = -548.104 kJ/mol – S^o(298.15 K) = 72.069 J/mol-K – p^o = 0.1 MPa



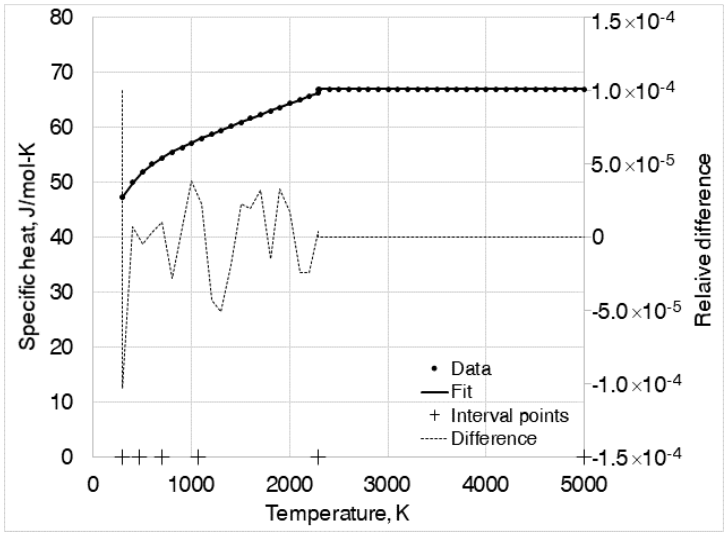


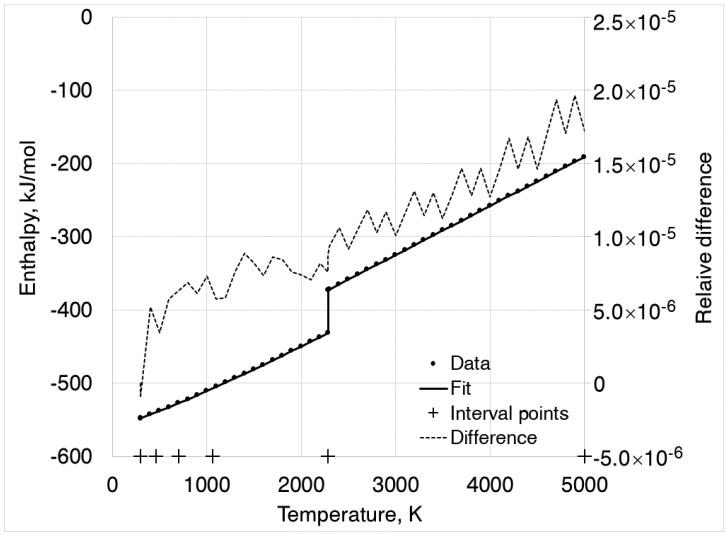


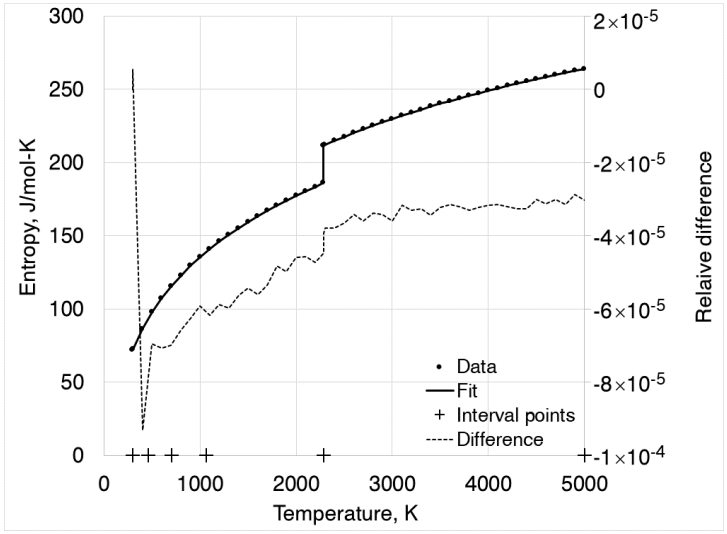


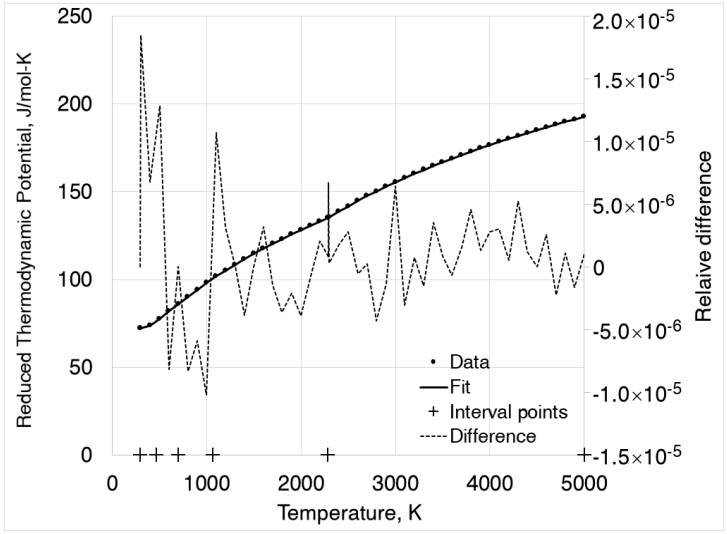
Barium Oxide (BaO) – Ba1O1(cr,l)

MW = 153.33 g/mol – $\Delta_f H^o$(298.15 K) = -548.104 kJ/mol – S^o(298.15 K) = 72.069 J/mol-K – p^o = 0.1 MPa



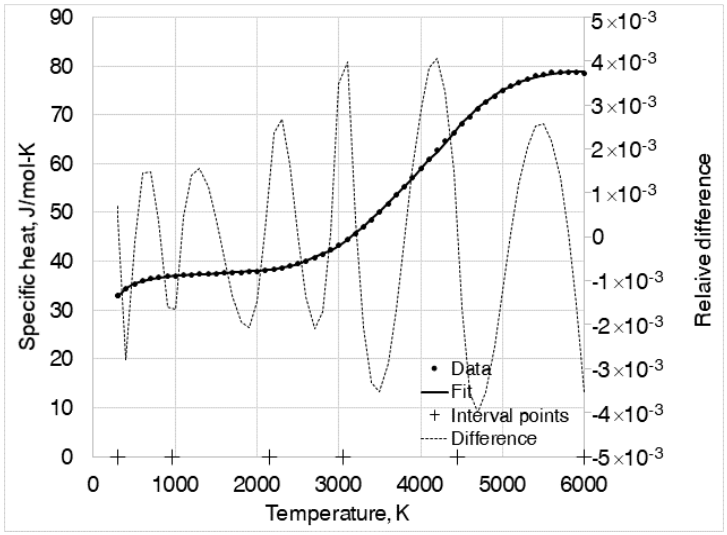


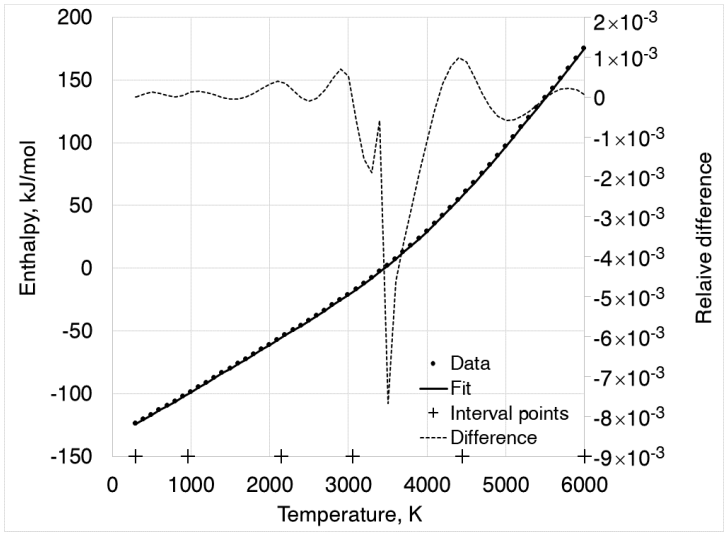


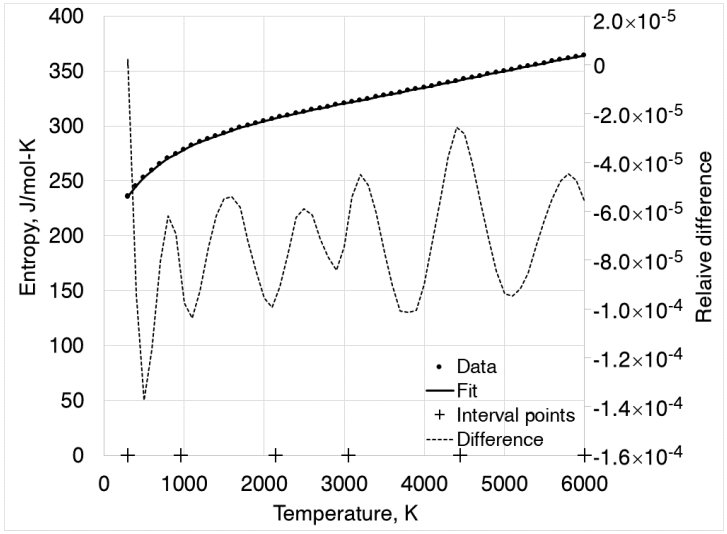


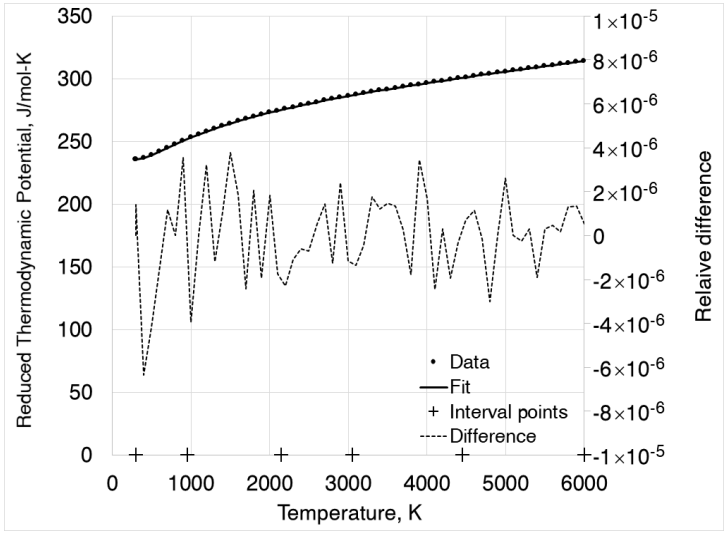
Barium Oxide (BaO) – Ba1O1(g)

MW = 153.33 g/mol – $\Delta_f H^o$(298.15 K) = -123.846 kJ/mol – S^o(298.15 K) = 235.457 J/mol-K – p^o = 0.1 MPa



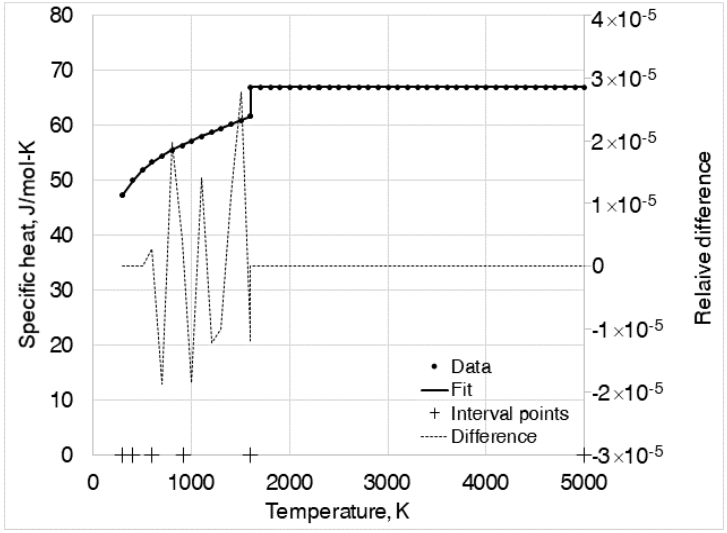


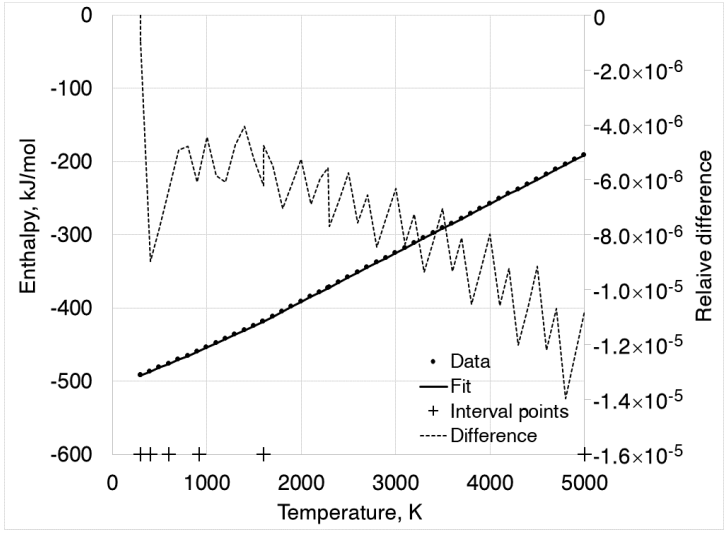


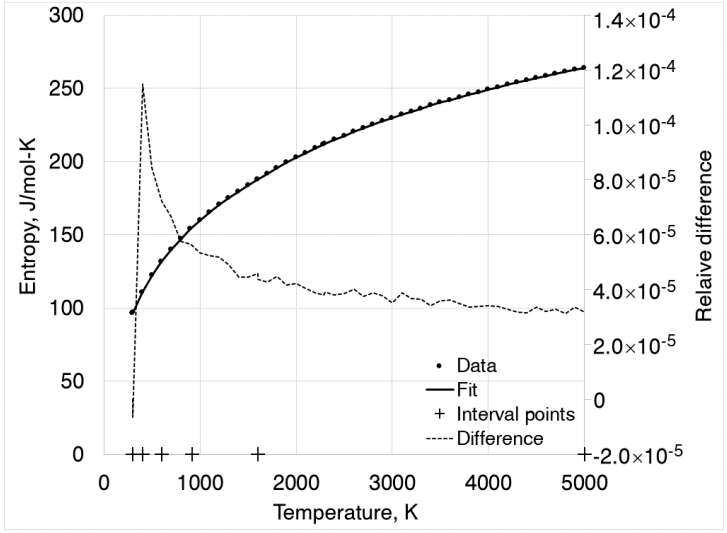


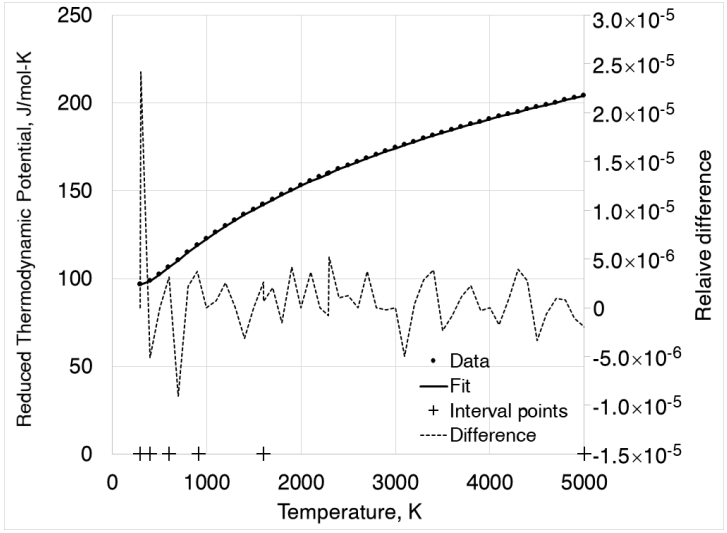
Barium Oxide (BaO) – Ba1O1(l)

MW = 153.33 g/mol – $\Delta_f H^o$(298.15 K) = -491.627 kJ/mol – S^o(298.15 K) = 96.553 J/mol-K – p^o = 0.1 MPa



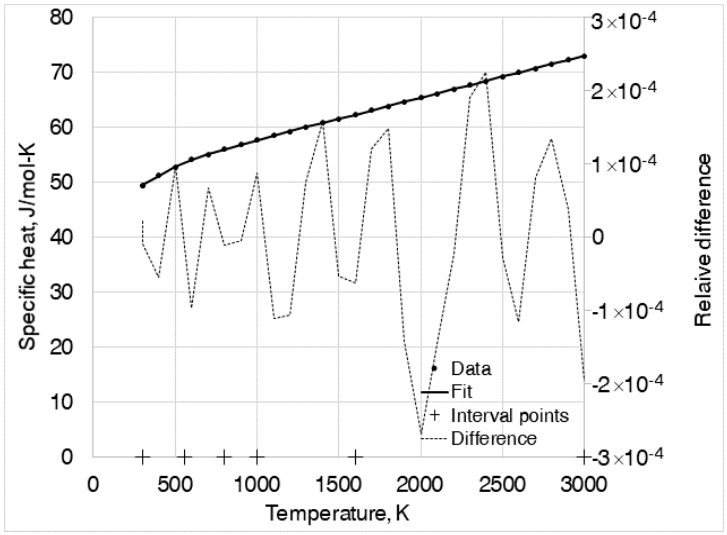


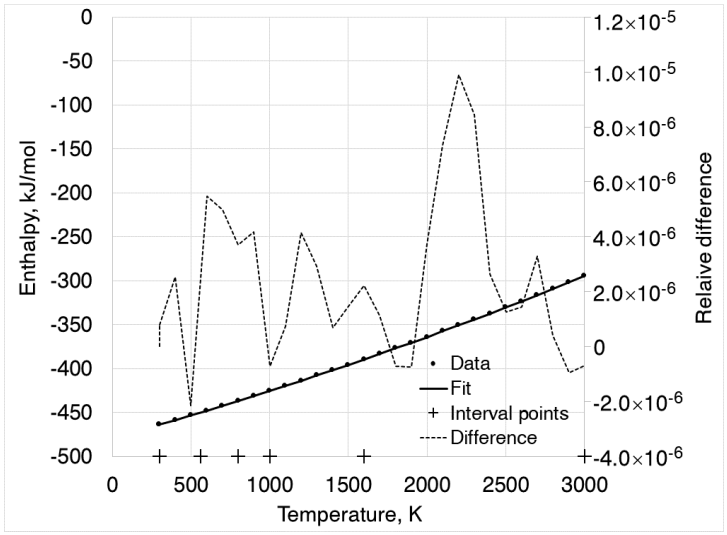


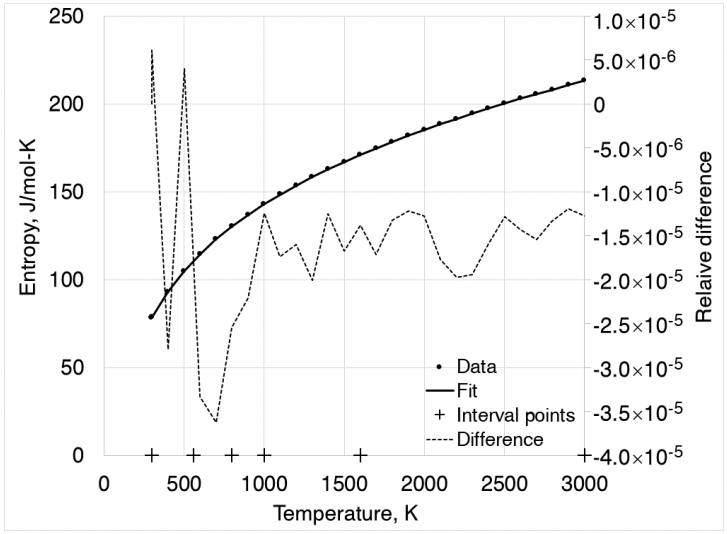


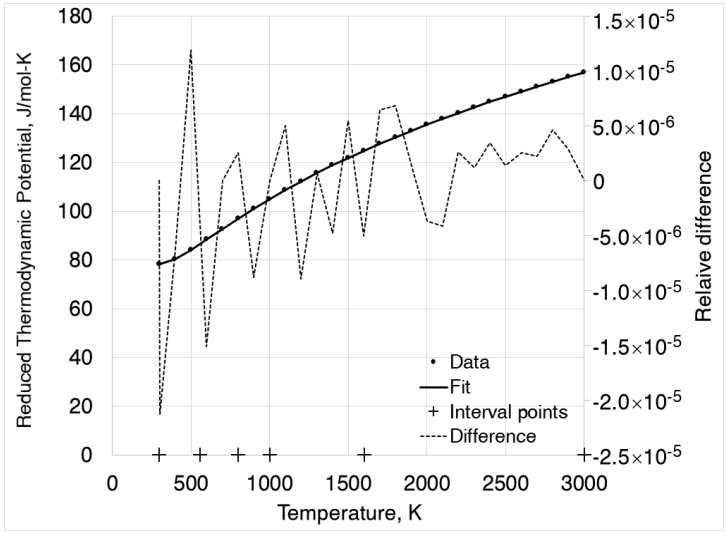
Barium Sulfide (BaS) – Ba1S1(cr)

MW = 169.39 g/mol – $\Delta_f H^o$(298.15 K) = -463.587 kJ/mol – S^o(298.15 K) = 78.395 J/mol-K – p^o = 0.1 MPa



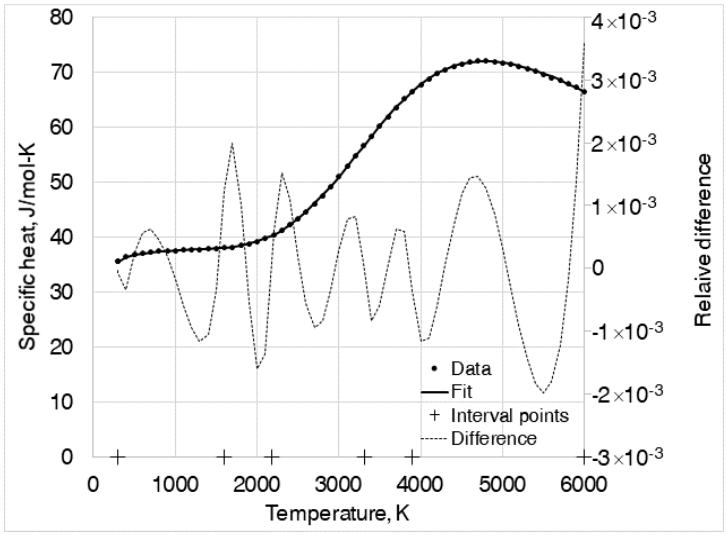


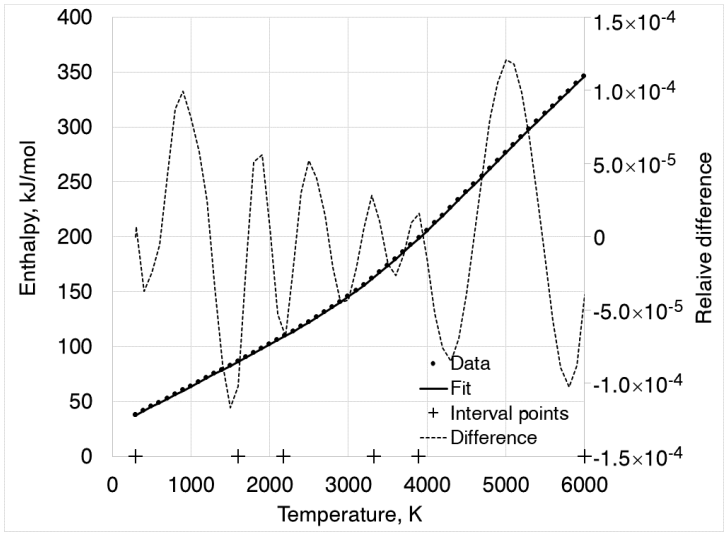


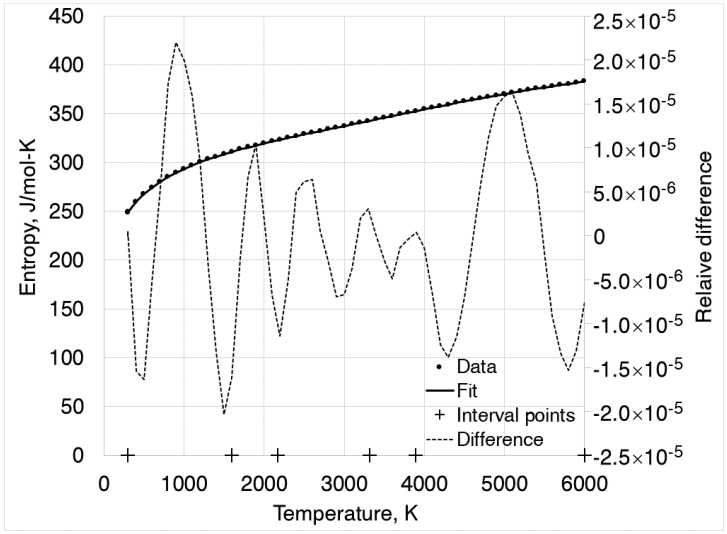


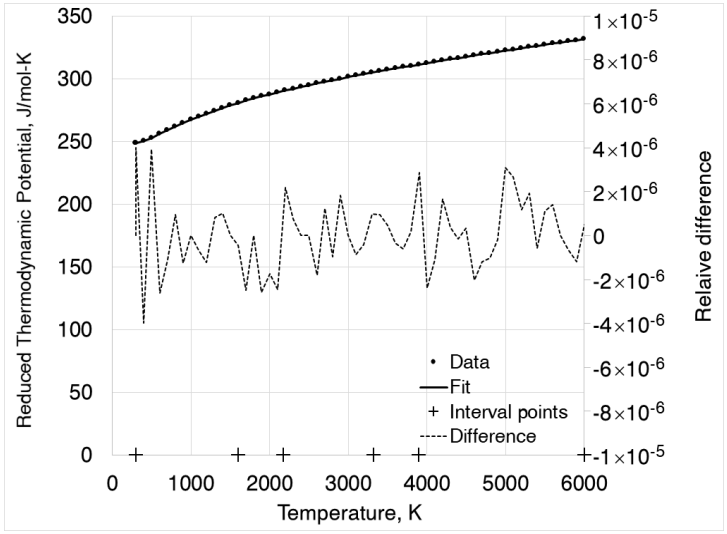
Barium Sulfide (BaS) – Ba1S1(g)

MW = 169.39 g/mol – $\Delta_f H^o$(298.15 K) = 37.782 kJ/mol – S^o(298.15 K) = 248.861 J/mol-K – p^o = 0.1 MPa



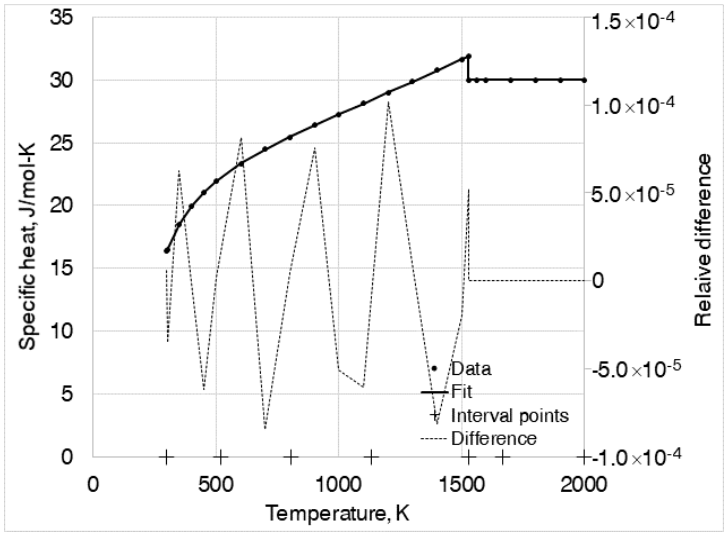


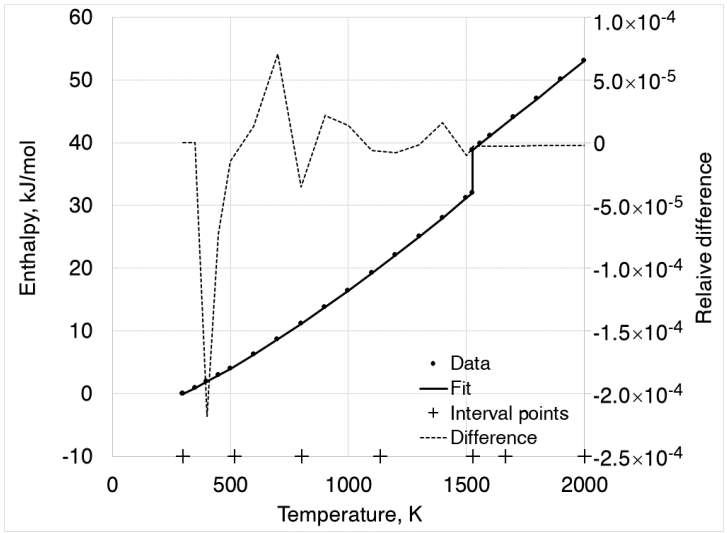


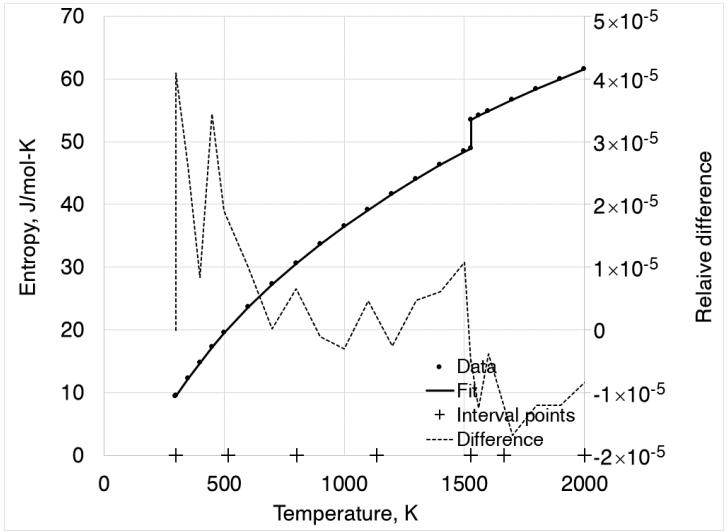


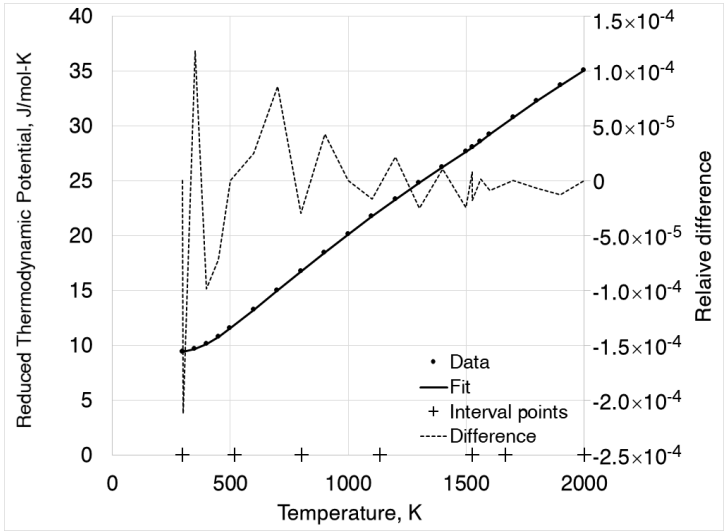
Beryllium (Be) – Be1(cr)

MW = 9.01218 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 9.44 J/mol-K – p^o = 0.1 MPa



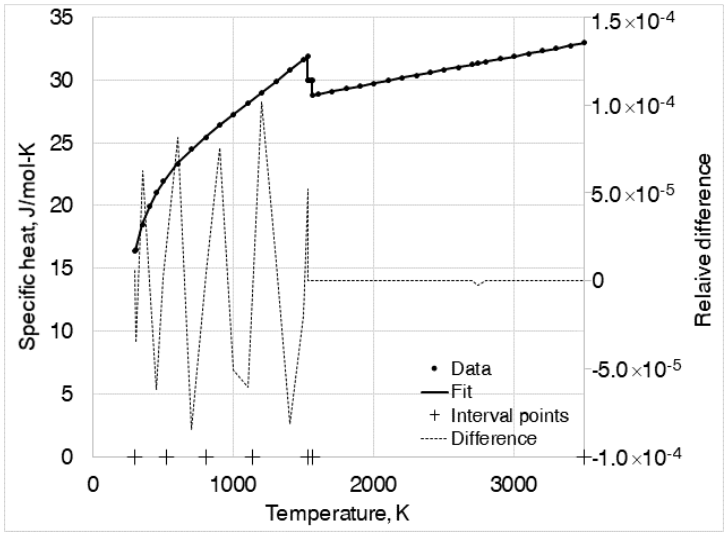


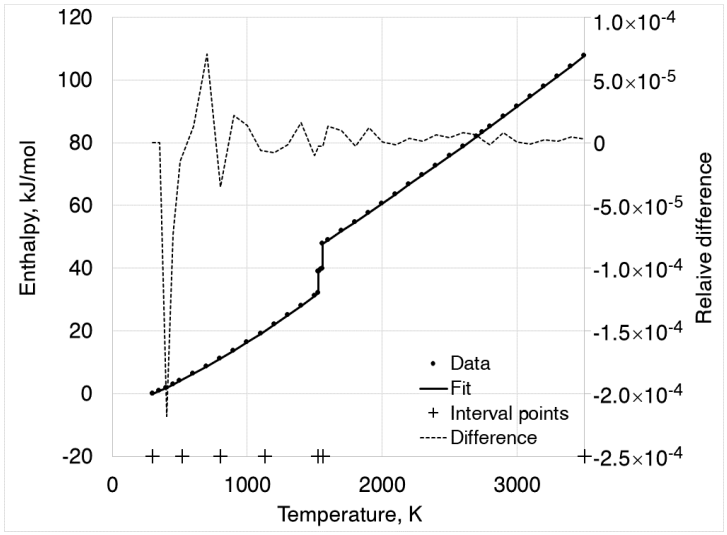


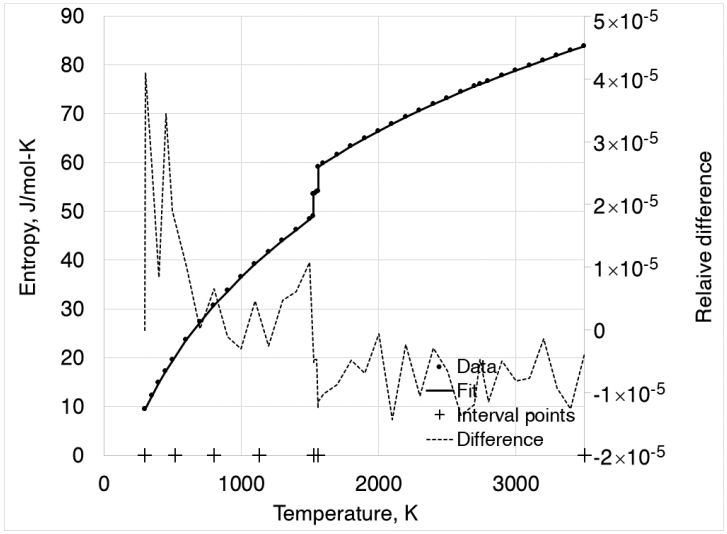


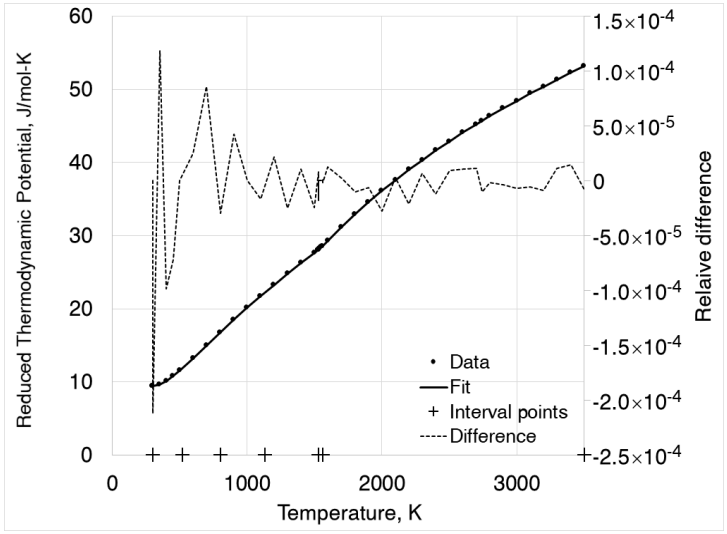
Beryllium (Be) – Be1(cr,l)

MW = 9.01218 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 9.44 J/mol-K – p^o = 0.1 MPa



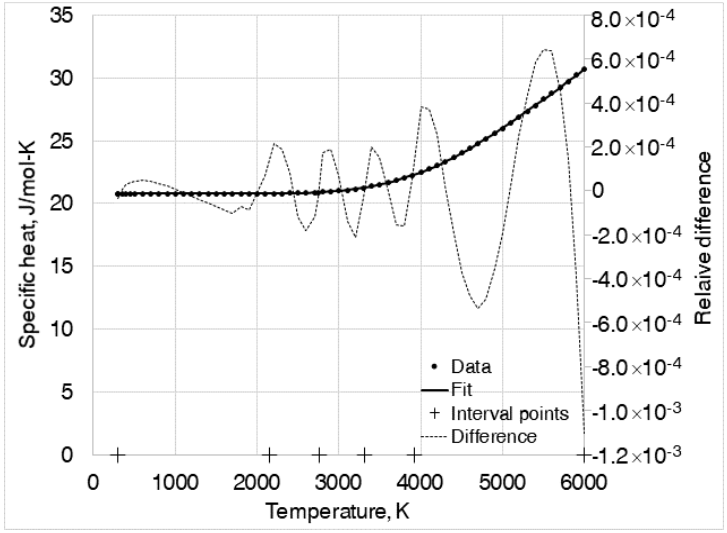


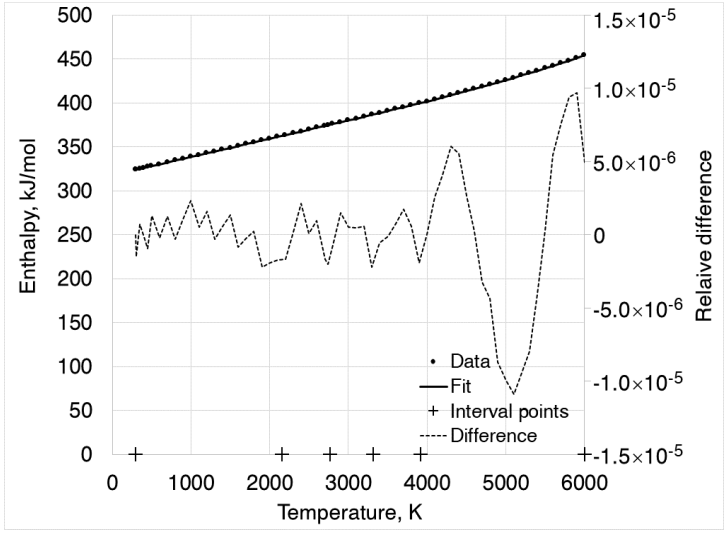


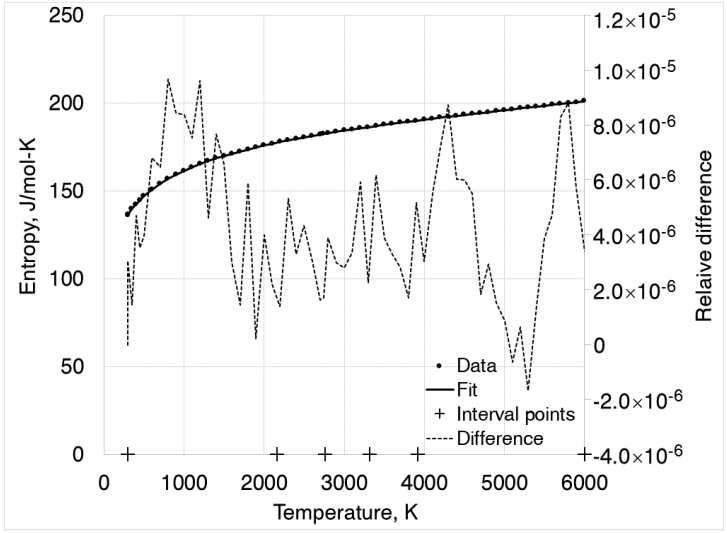


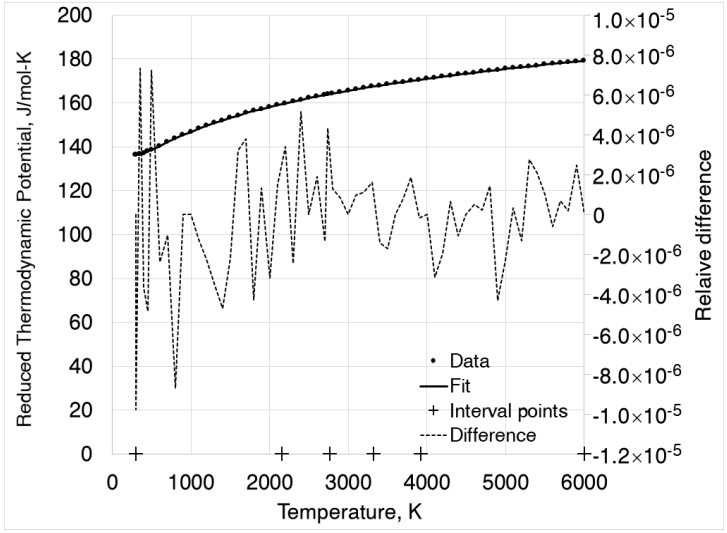
Beryllium (Be) – Be1(g)

MW = 9.01218 g/mol – $\Delta_f H^o$(298.15 K) = 324 kJ/mol – S^o(298.15 K) = 136.274 J/mol-K – p^o = 0.1 MPa



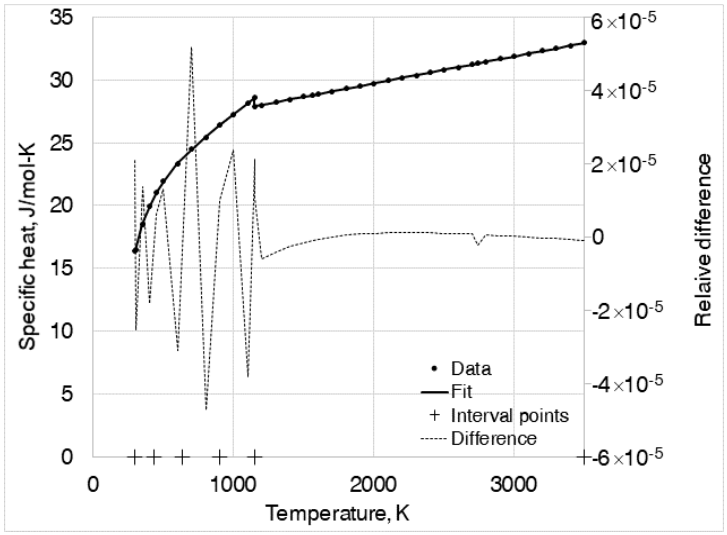


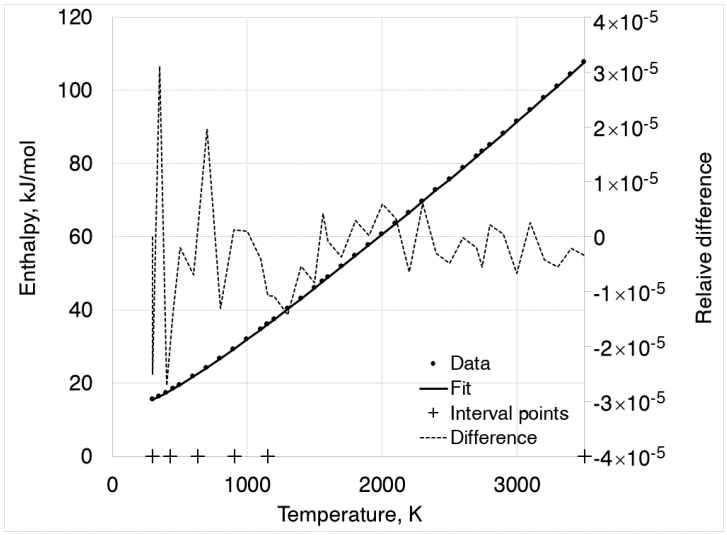


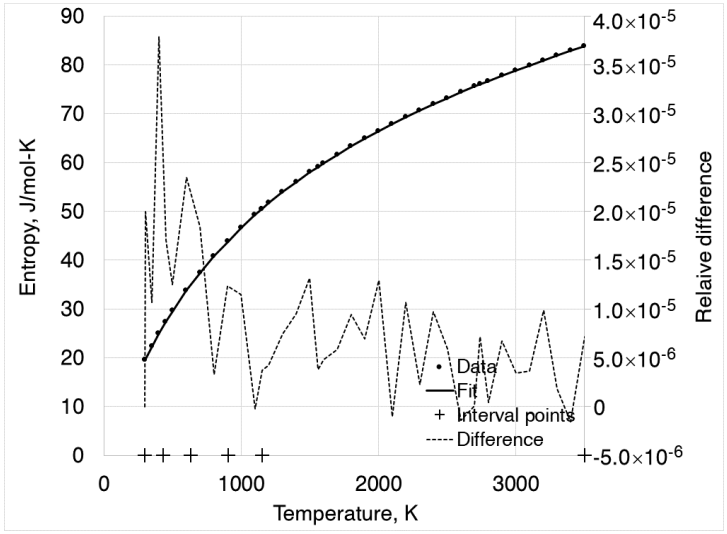


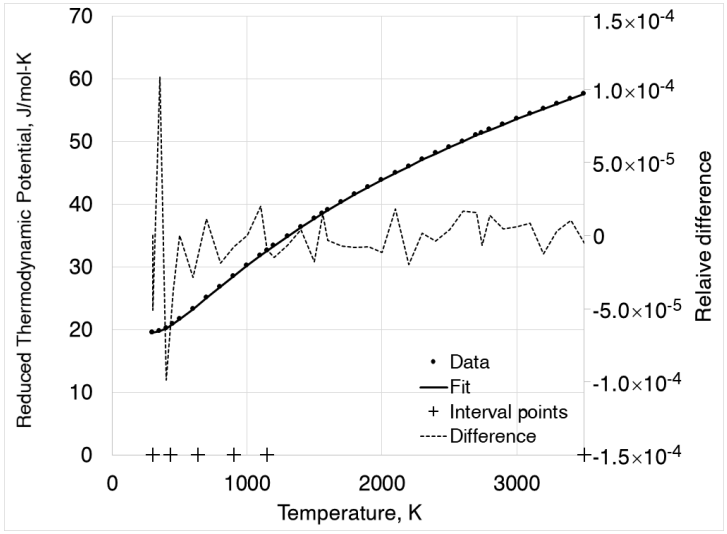
Beryllium (Be) – Be1(l)

MW = 9.01218 g/mol – $\Delta_f H^o$(298.15 K) = 15.508 kJ/mol – S^o(298.15 K) = 19.539 J/mol-K – p^o = 0.1 MPa



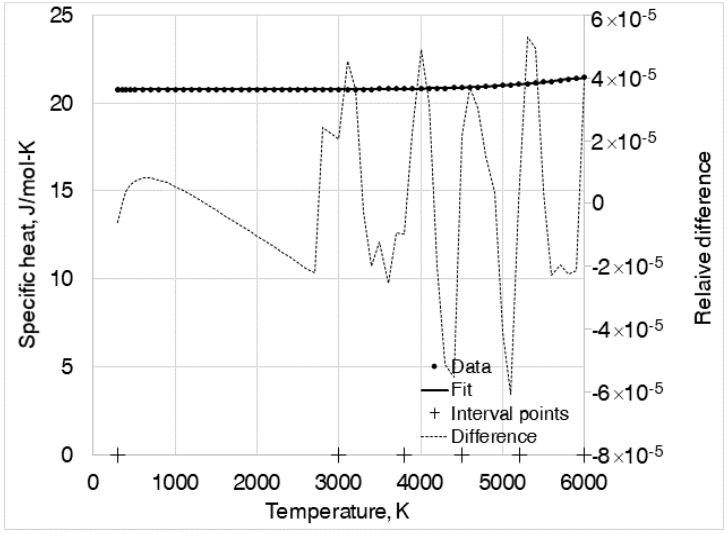


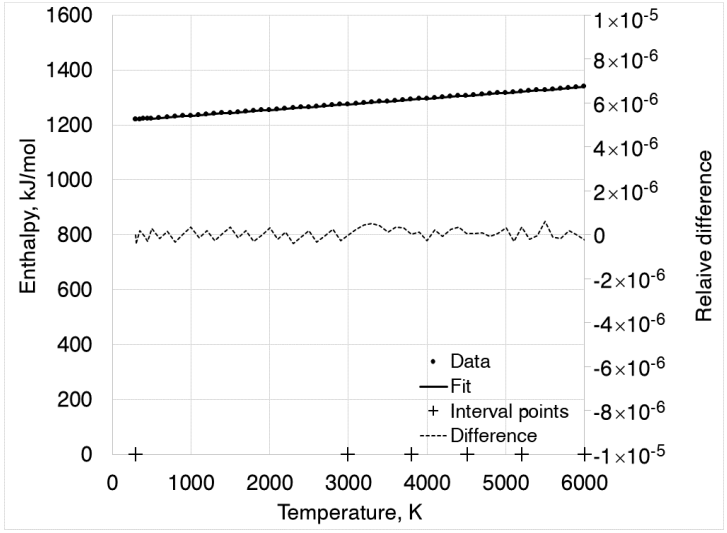


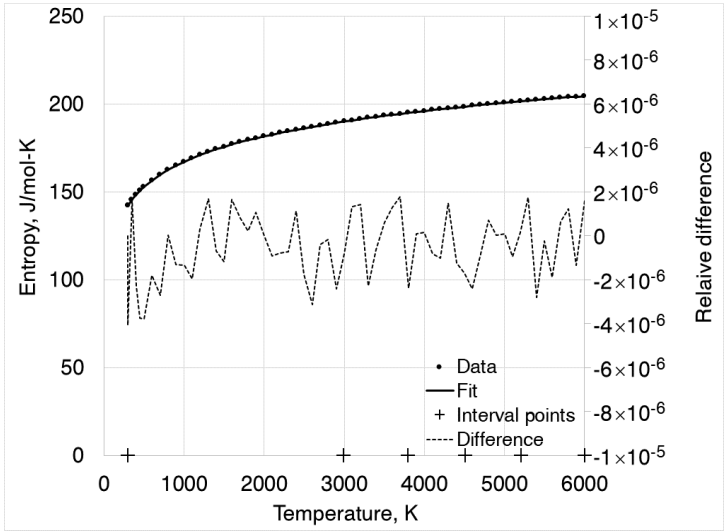


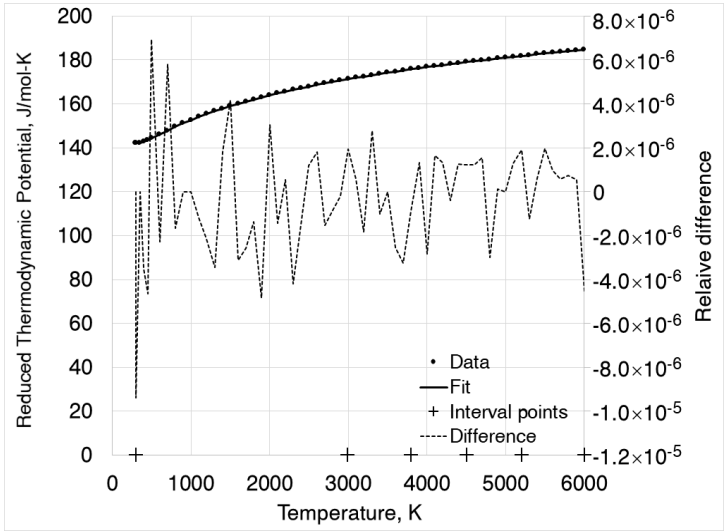
Beryllium, Ion (Be^+) – Be1+(g)

MW = 9.01163 g/mol – $\Delta_f H^o$(298.15 K) = 1219.232 kJ/mol – S^o(298.15 K) = 142.037 J/mol-K – p^o = 0.1 MPa



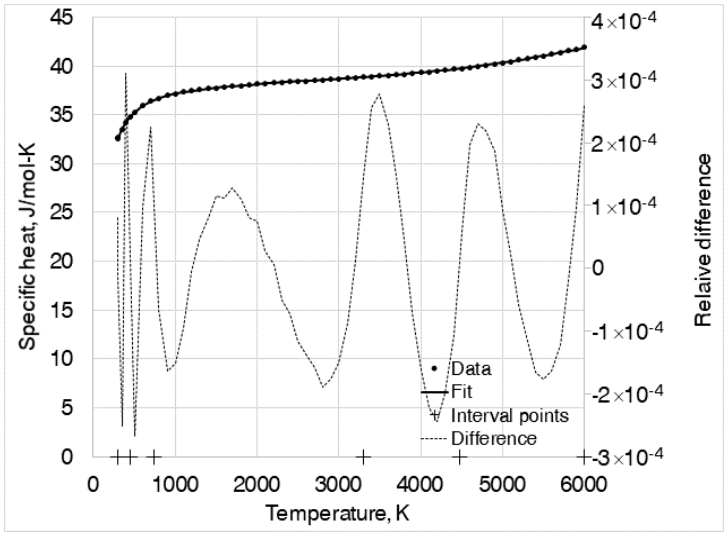


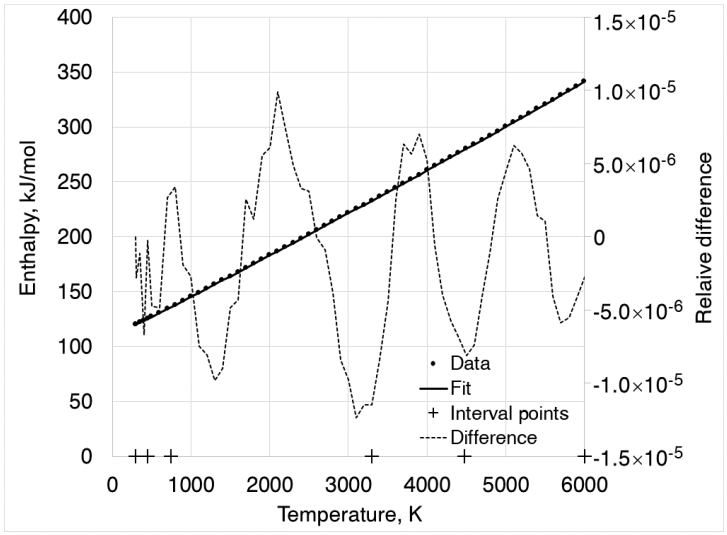


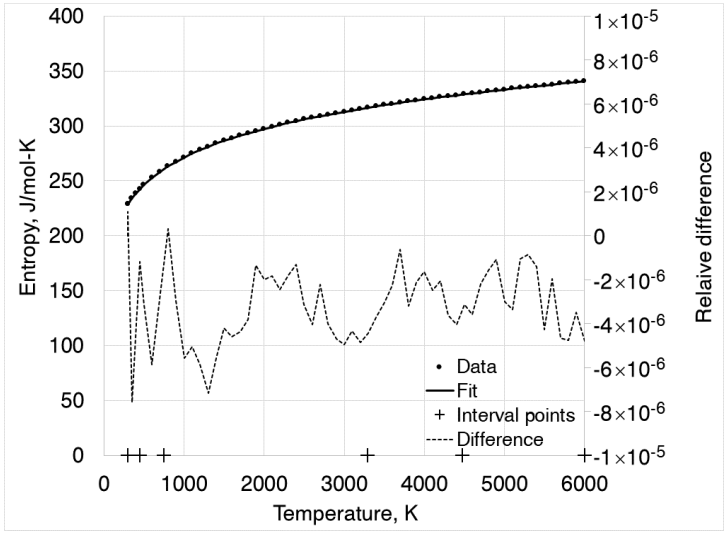


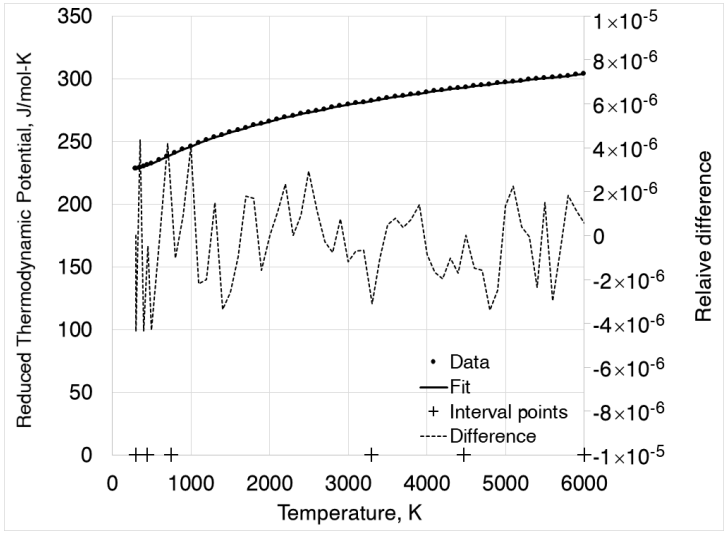
Beryllium Bromide (BeBr) – Be1Br1(g)

MW = 88.916 g/mol – $\Delta_f H^o$(298.15 K) = 120.11 kJ/mol – S^o(298.15 K) = 228.534 J/mol-K – p^o = 0.1 MPa



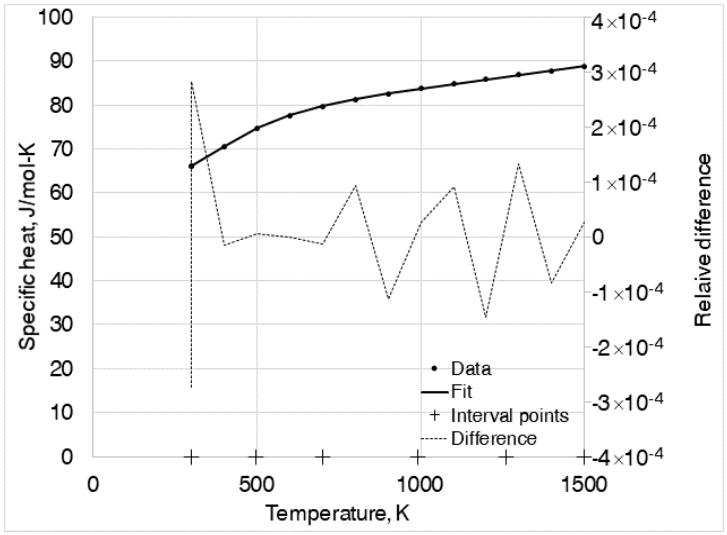


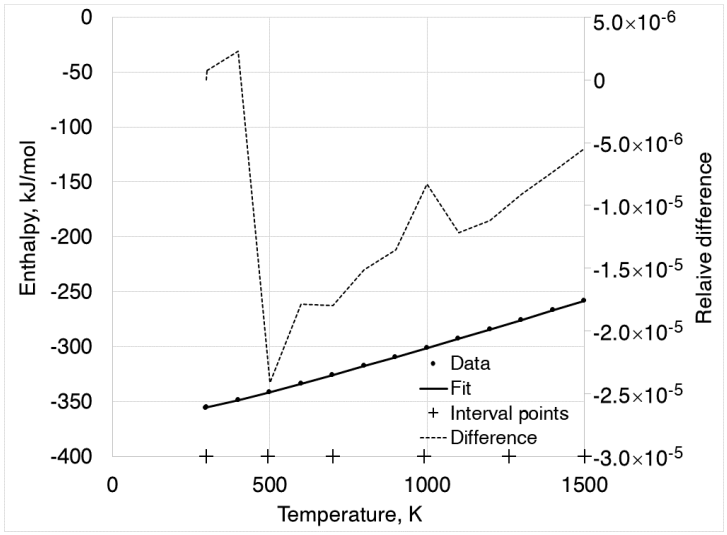


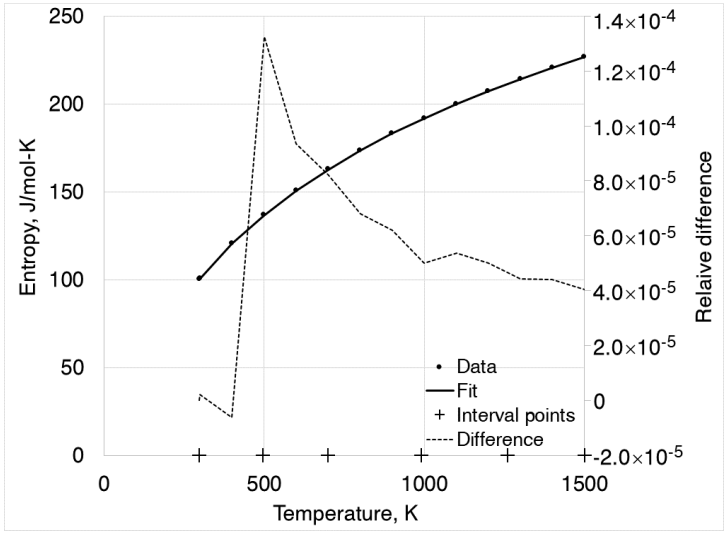


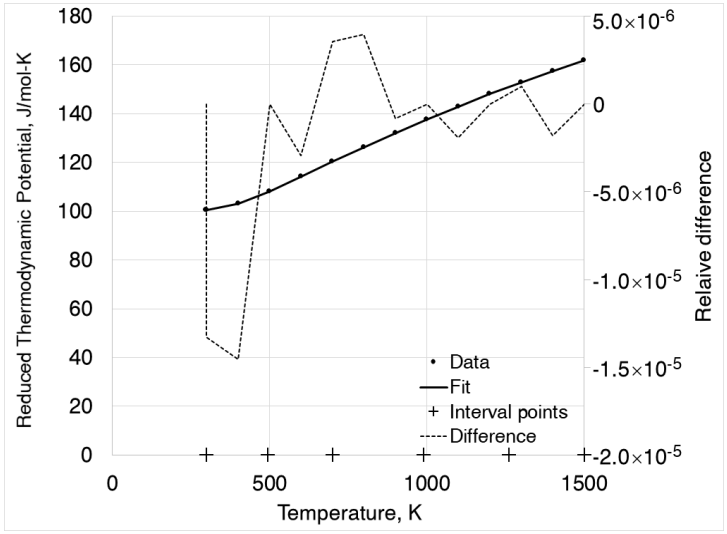
Beryllium Bromide ($BeBr_2$) – Be1Br2(cr)

MW = 168.82 g/mol – $\Delta_f H^o$(298.15 K) = -355.64 kJ/mol – S^o(298.15 K) = 100.416 J/mol-K – p^o = 0.1 MPa



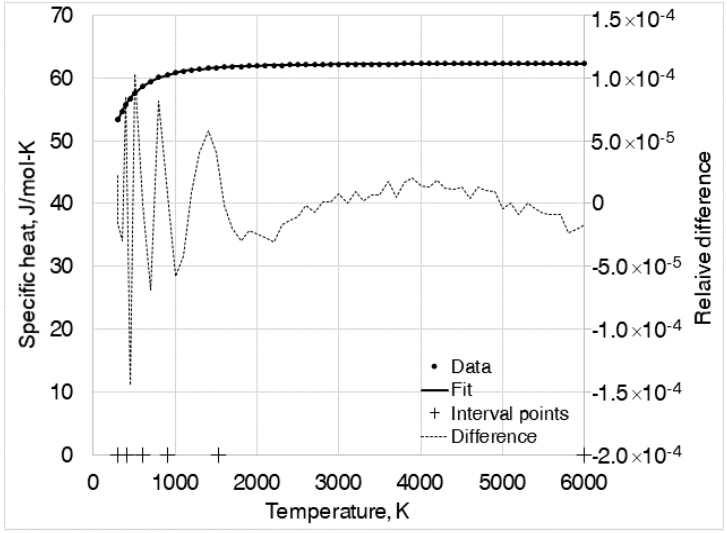


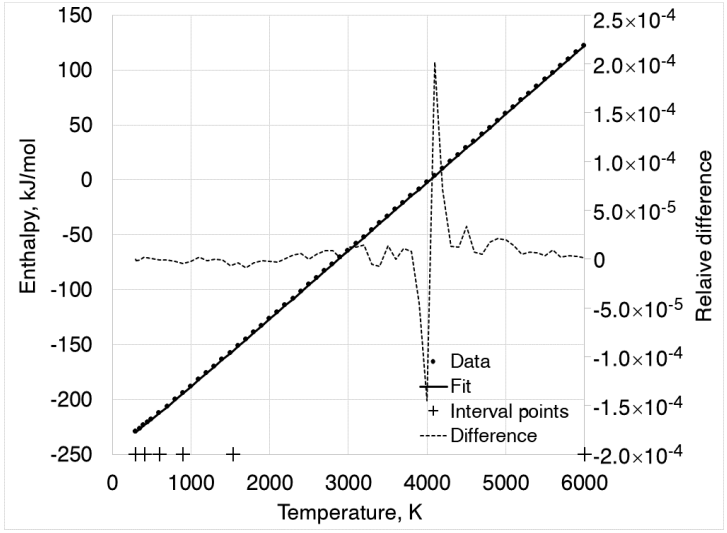


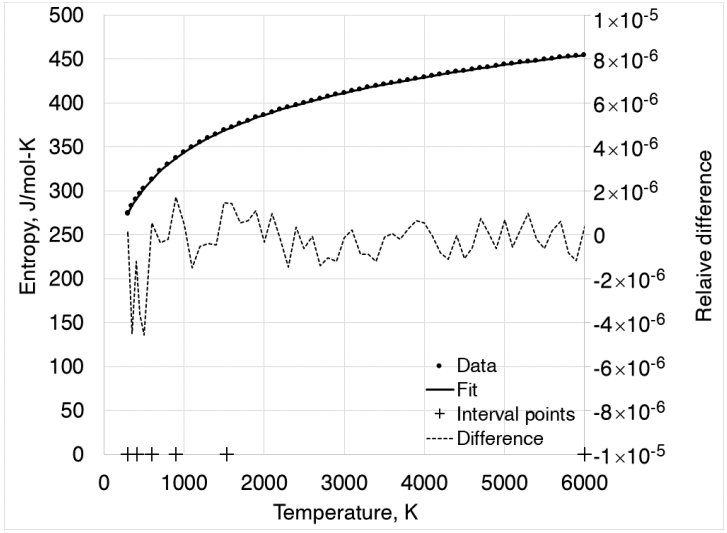


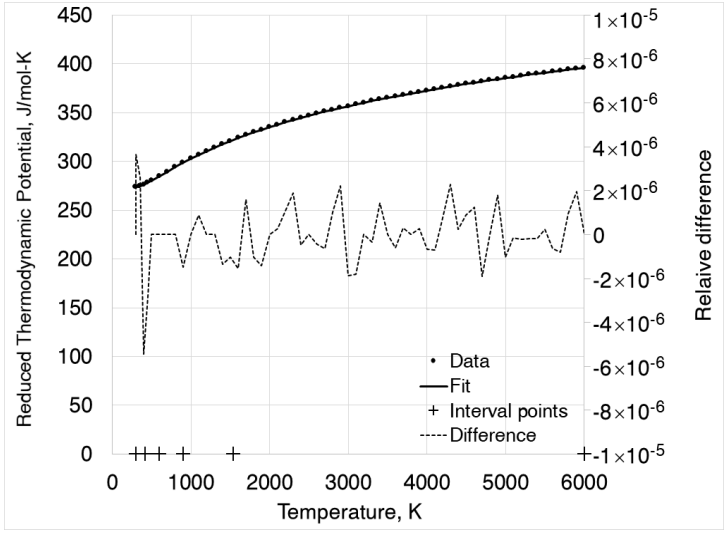
Beryllium Bromide ($BeBr_2$) – Be1Br2(g)

MW = 168.82 g/mol – $\Delta_f H^o$(298.15 K) = -229.283 kJ/mol – S^o(298.15 K) = 273.909 J/mol-K – p^o = 0.1 MPa



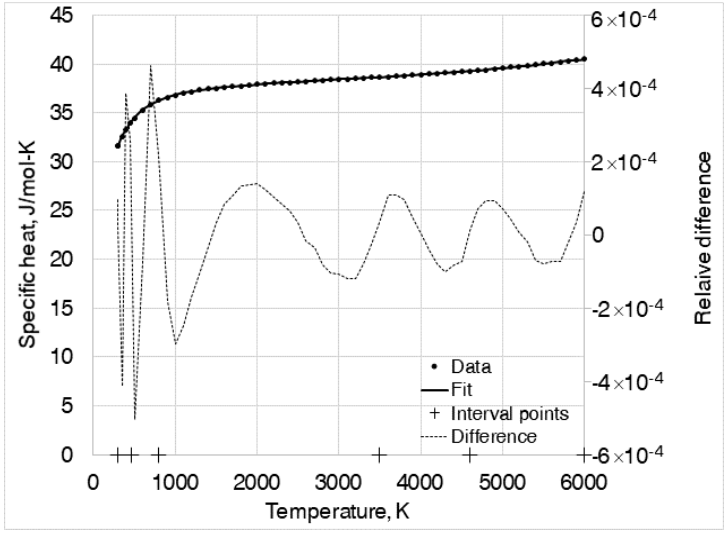


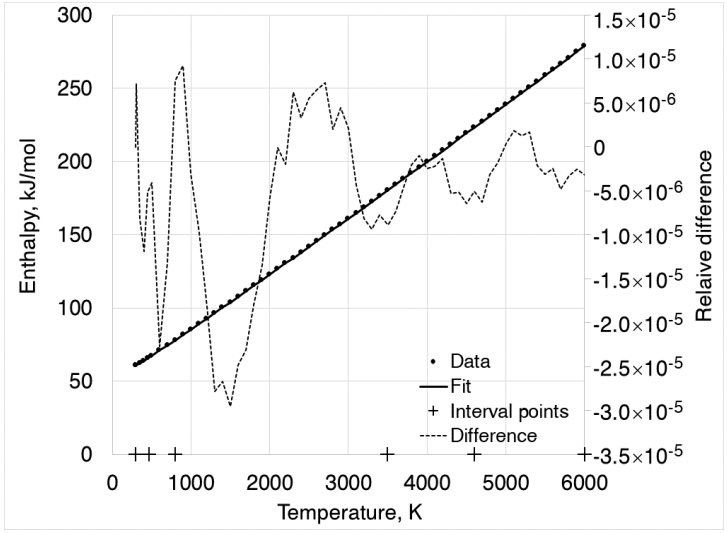


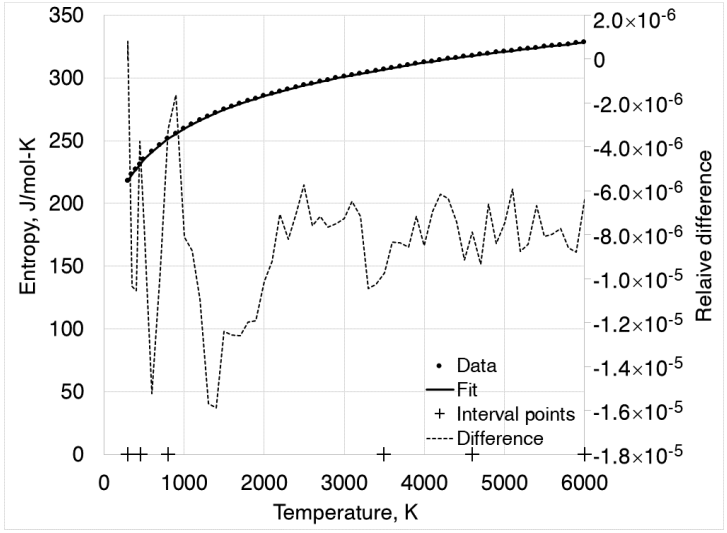


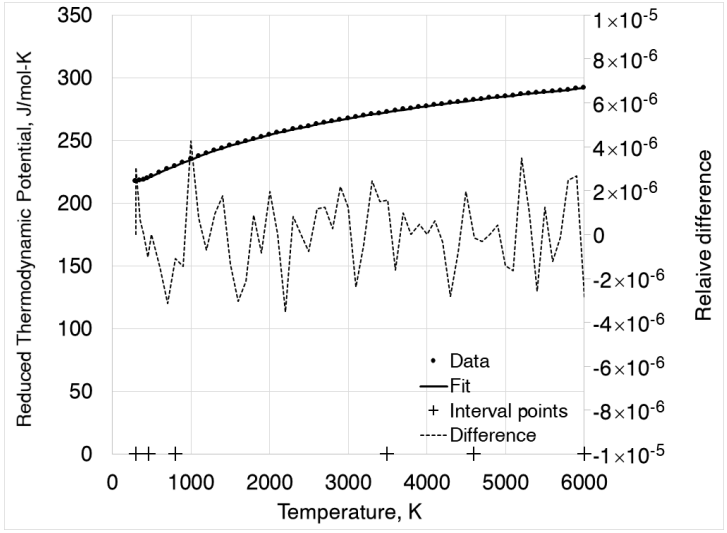
Beryllium Chloride (BeCl) – Be1Cl1(g)

MW = 44.465 g/mol – $\Delta_f H^o$(298.15 K) = 60.668 kJ/mol – S^o(298.15 K) = 217.615 J/mol-K – p^o = 0.1 MPa



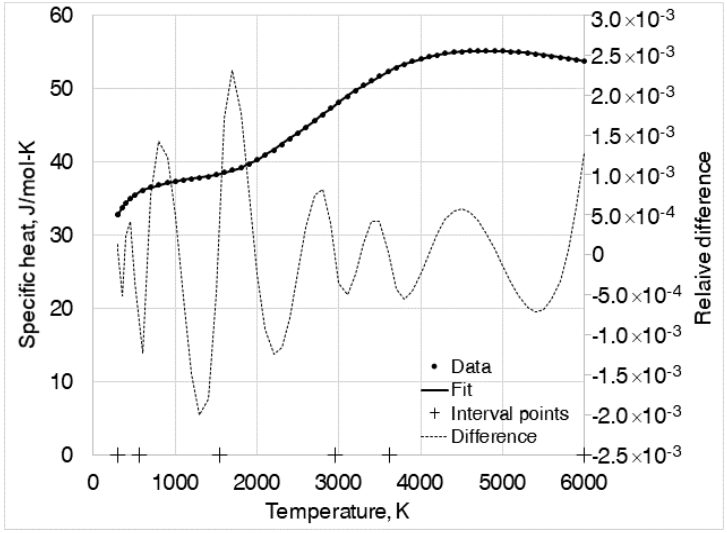


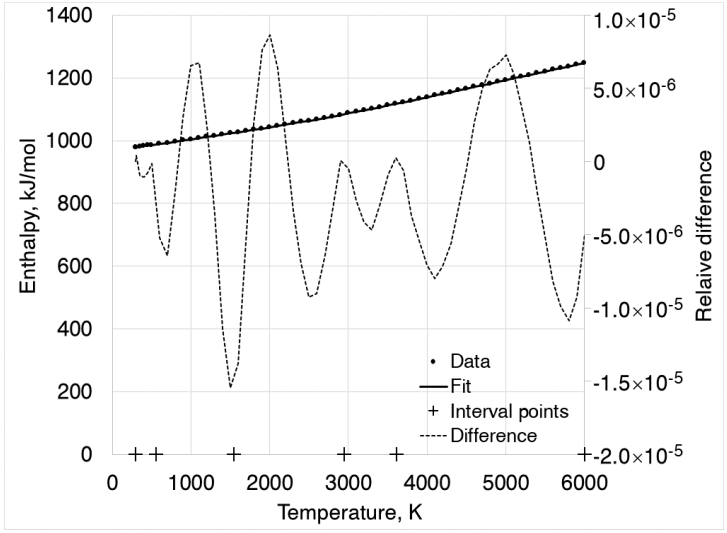


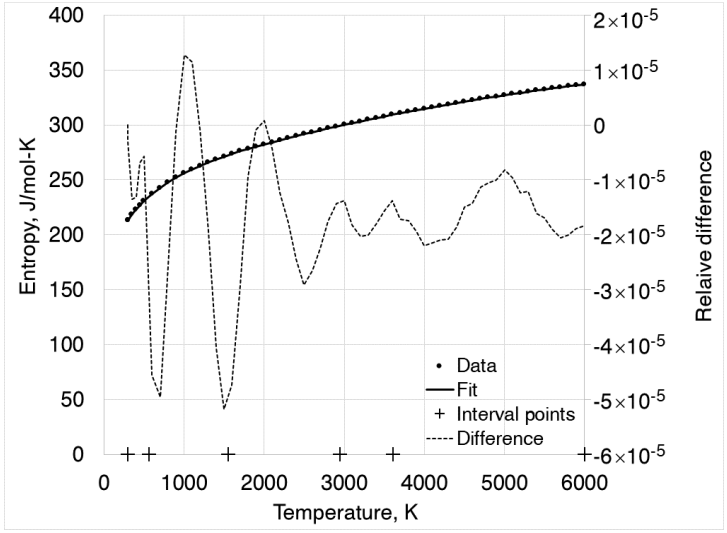


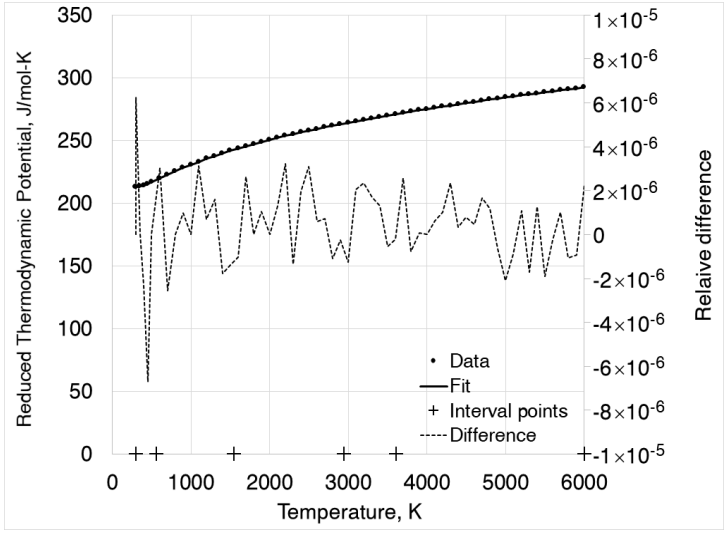
Beryllium Chloride, Ion ($BeCl^+$) – Be1Cl1+(g)

MW = 44.465 g/mol – $\Delta_f H^o$(298.15 K) = 979.056 kJ/mol – S^o(298.15 K) = 212.955 J/mol-K – p^o = 0.1 MPa



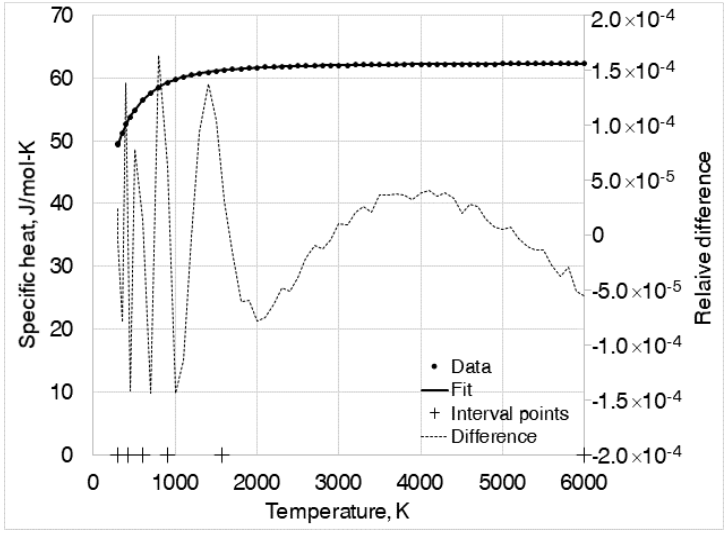


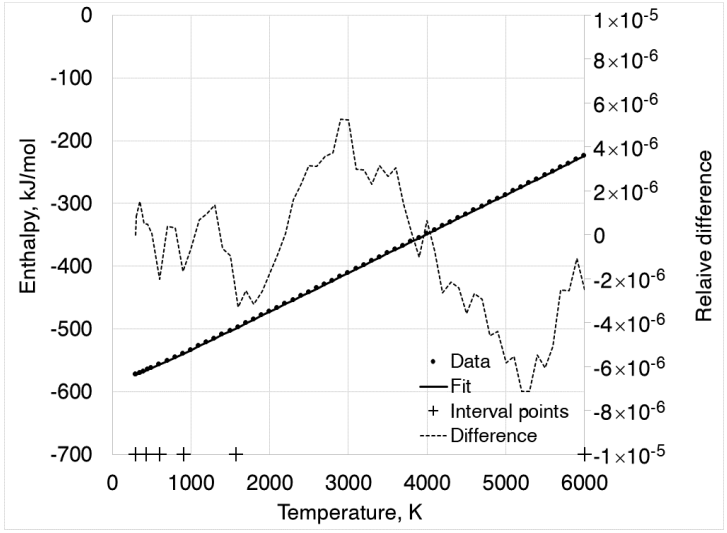


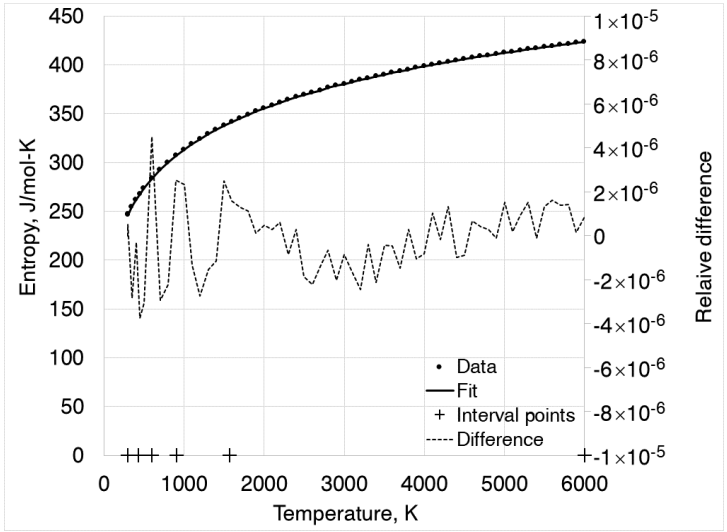


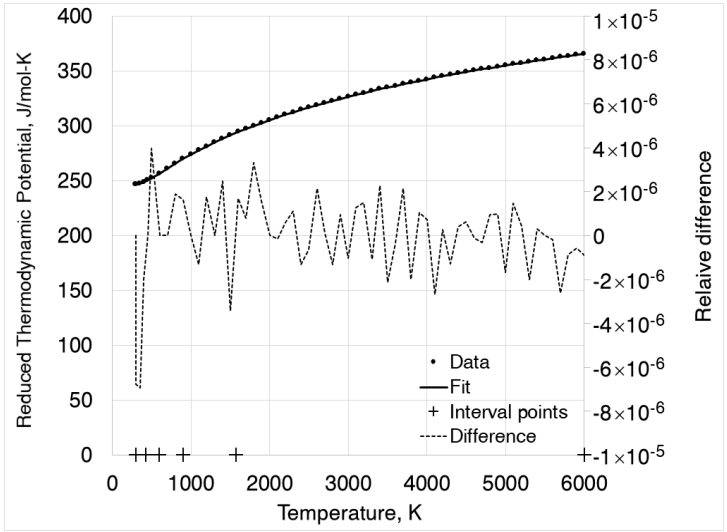
Beryllium Chloride Fluoride (BeClF) – Be1Cl1F1(g)

MW = 63.464 g/mol – $\Delta_f H^o$(298.15 K) = -573.208 kJ/mol – S^o(298.15 K) = 246.486 J/mol-K – p^o = 0.1 MPa



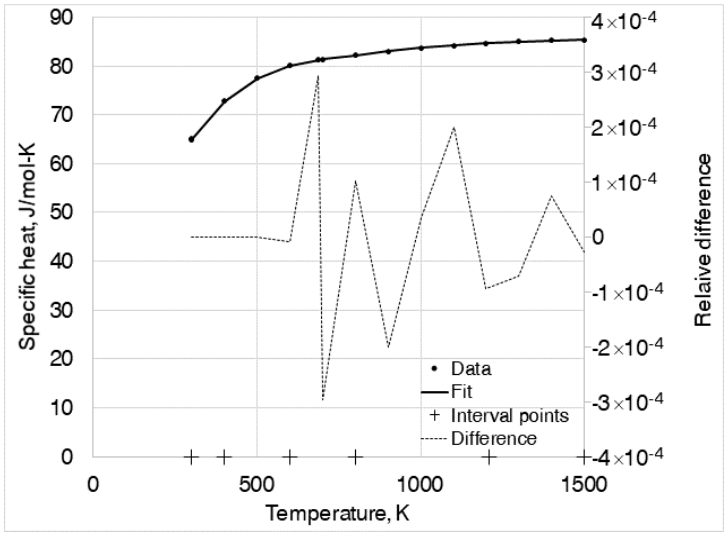


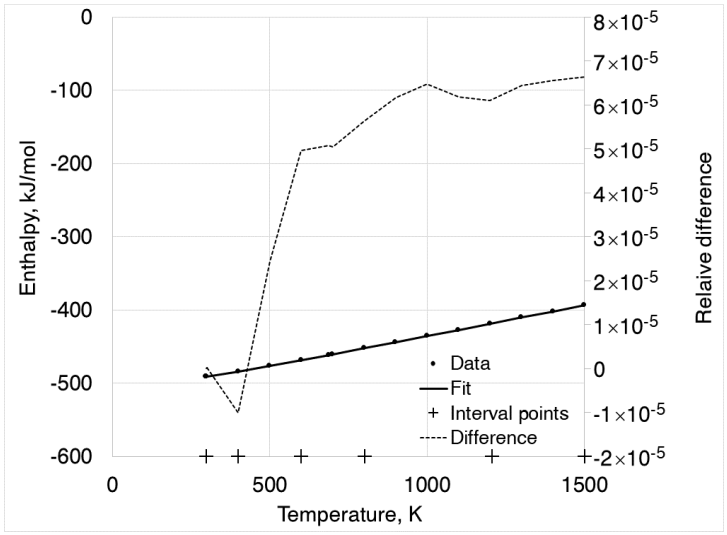


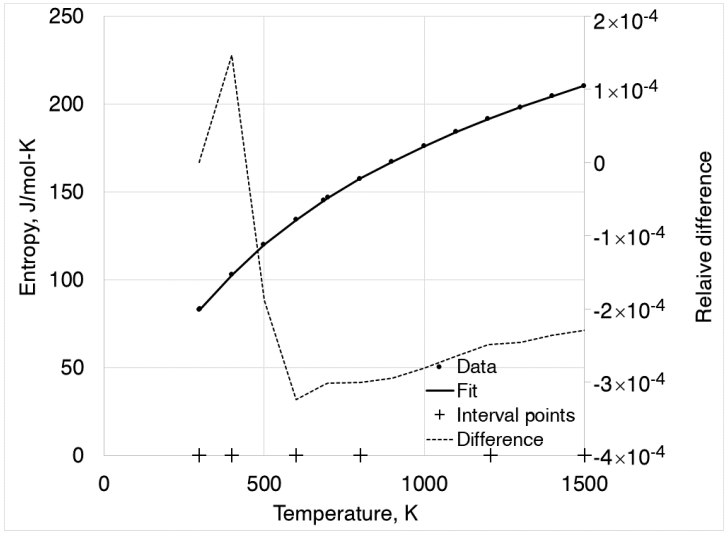


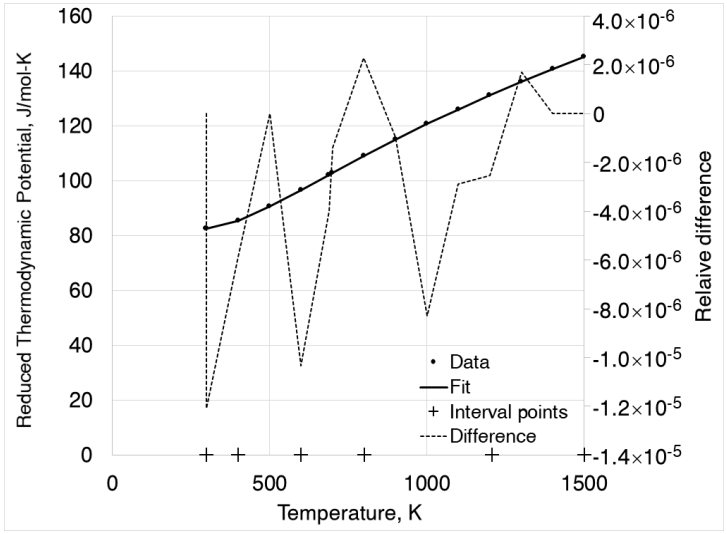
Beryllium Chloride, Alpha ($BeCl_2$) – Be1Cl2(cr)_01

MW = 79.918 g/mol – $\Delta_f H^o$(298.15 K) = -490.93 kJ/mol – S^o(298.15 K) = 82.676 J/mol-K – p^o = 0.1 MPa



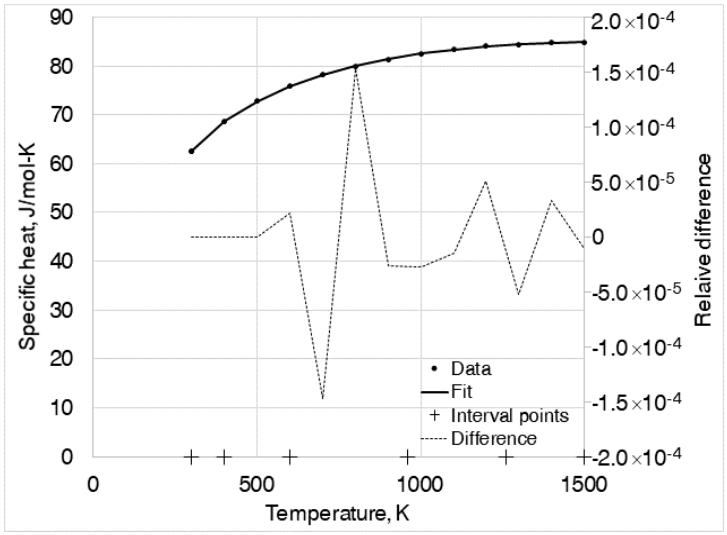


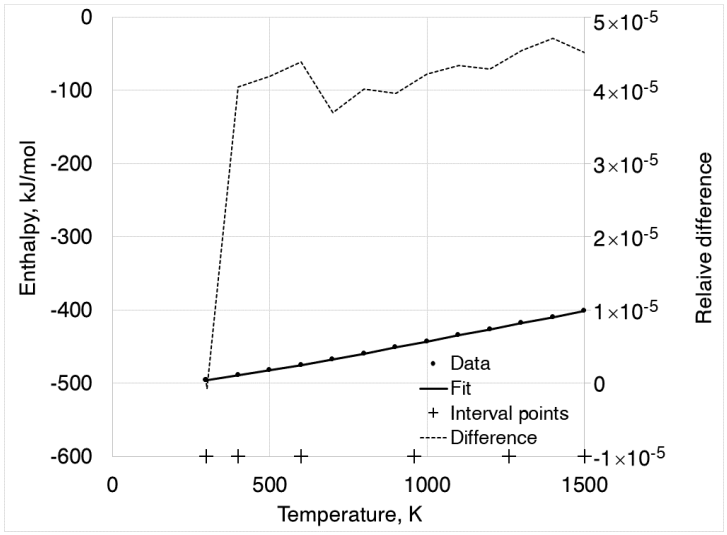


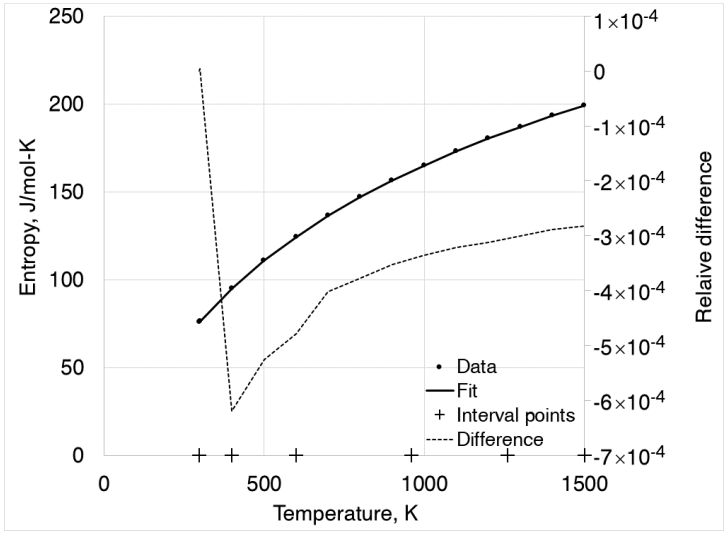


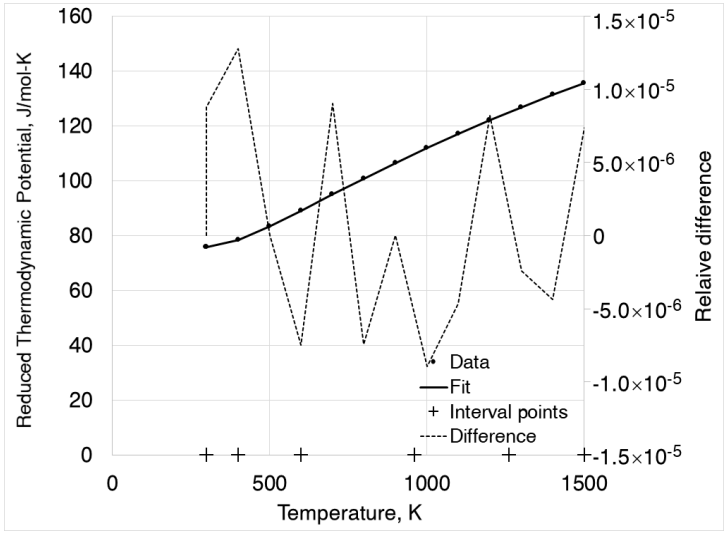
Beryllium Chloride, Beta ($BeCl_2$) – Be1Cl2(cr)_02

MW = 79.918 g/mol – $\Delta_f H^o$(298.15 K) = -496.222 kJ/mol – S^o(298.15 K) = 75.814 J/mol-K – p^o = 0.1 MPa



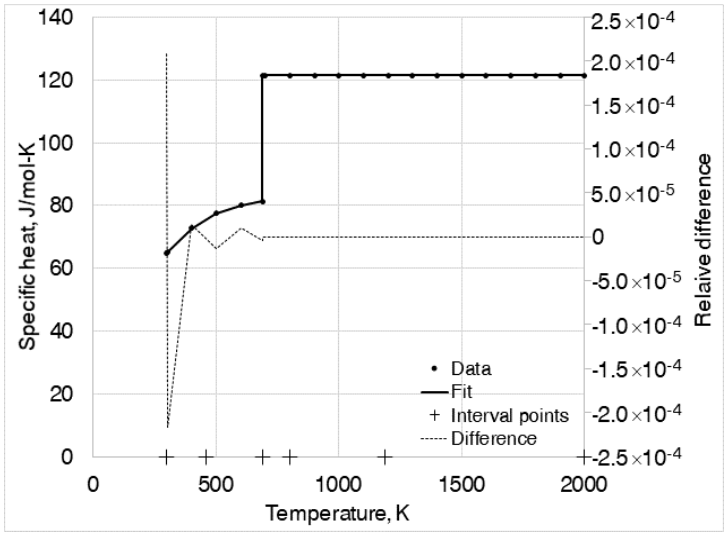


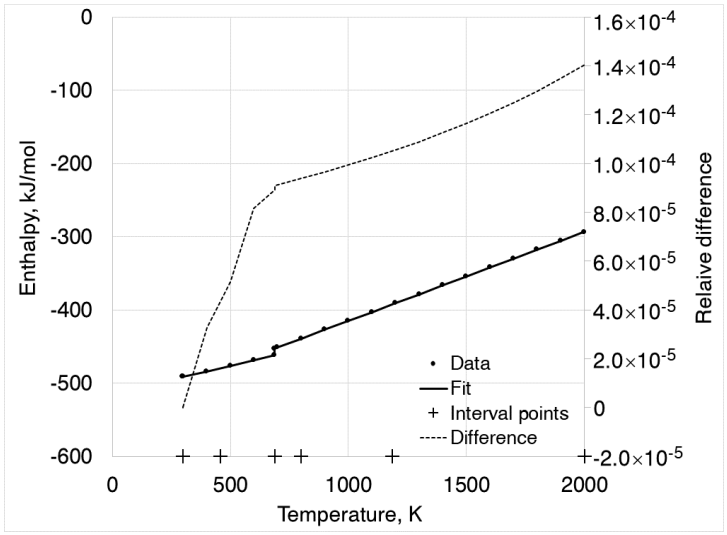


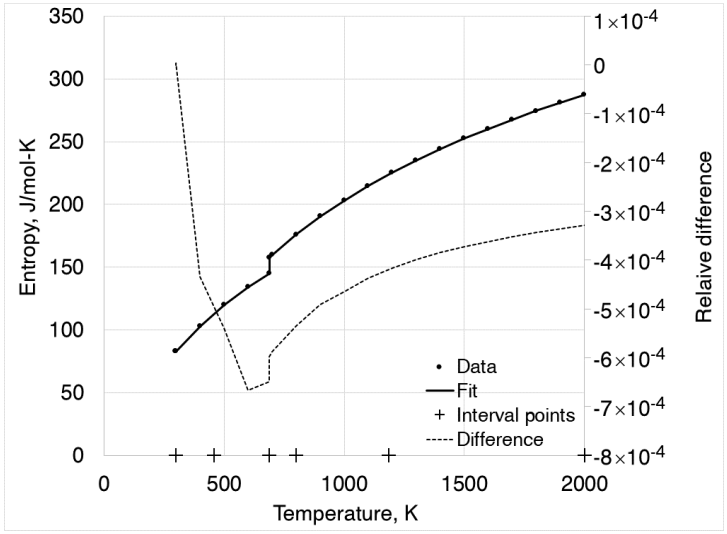


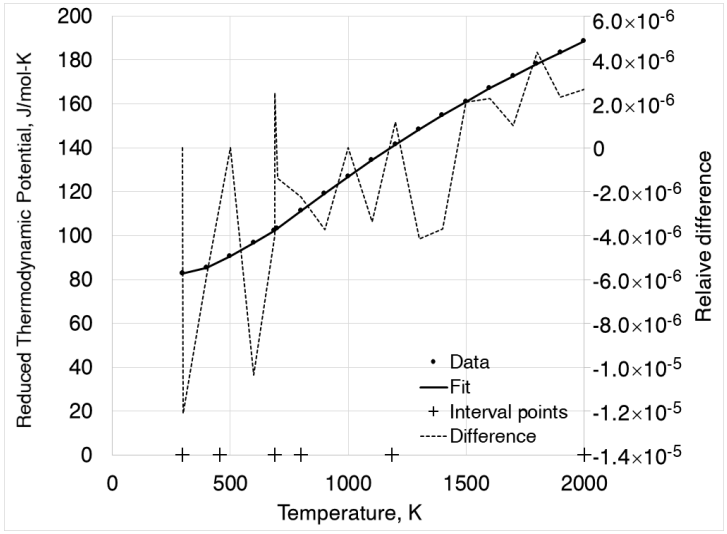
Beryllium Chloride ($BeCl_2$) – Be1Cl2(cr,l)

MW = 79.918 g/mol – $\Delta_f H^o$(298.15 K) = -490.93 kJ/mol – S^o(298.15 K) = 82.676 J/mol-K – p^o = 0.1 MPa



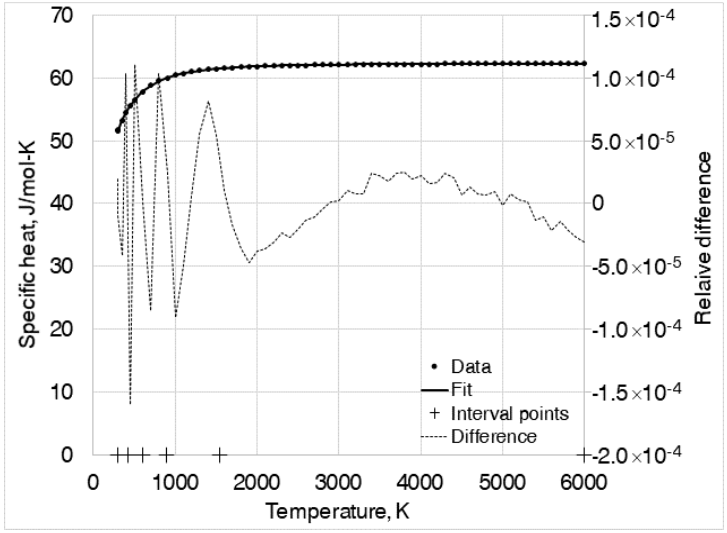


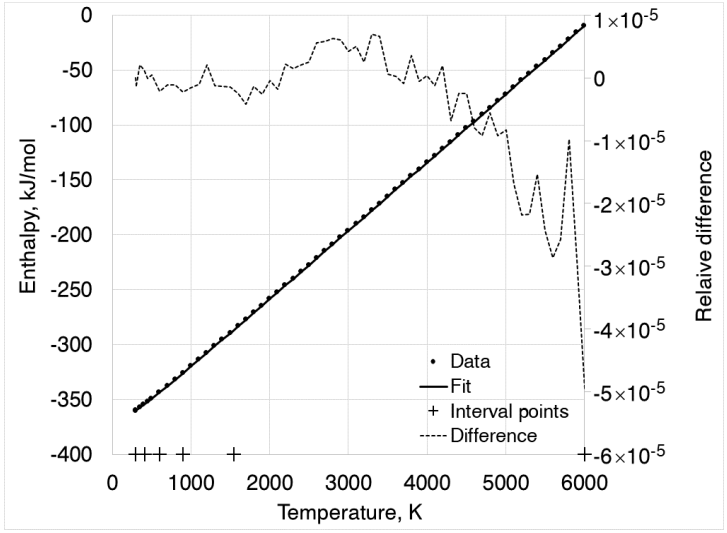


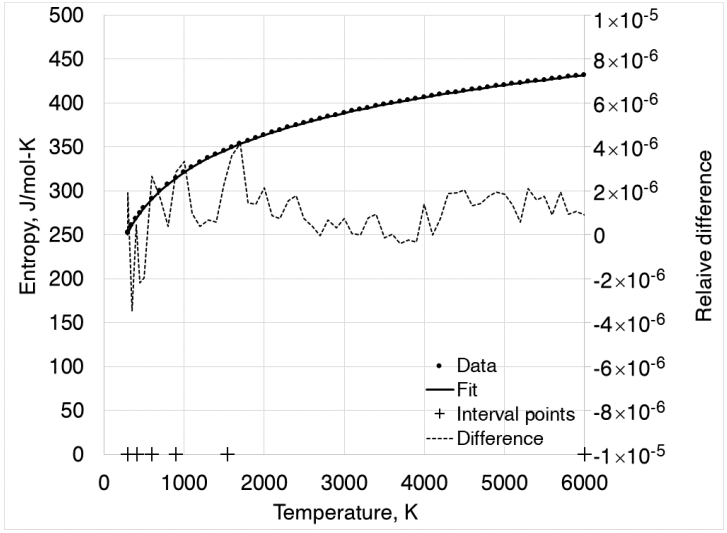


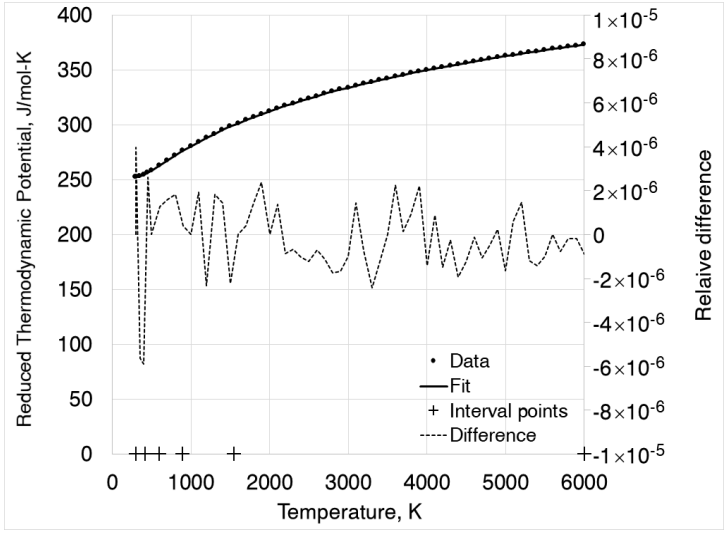
Beryllium Chloride ($BeCl_2$) – Be1Cl2(g)

MW = 79.918 g/mol – $\Delta_f H^o$(298.15 K) = -360.242 kJ/mol – S^o(298.15 K) = 252.232 J/mol-K – p^o = 0.1 MPa



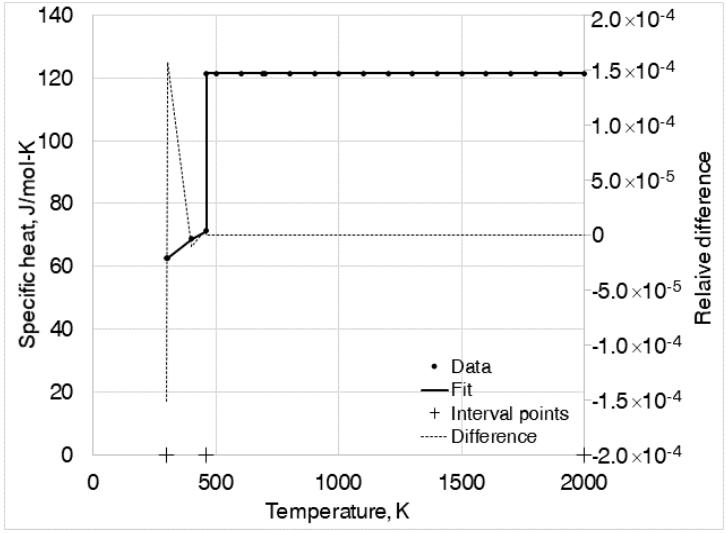


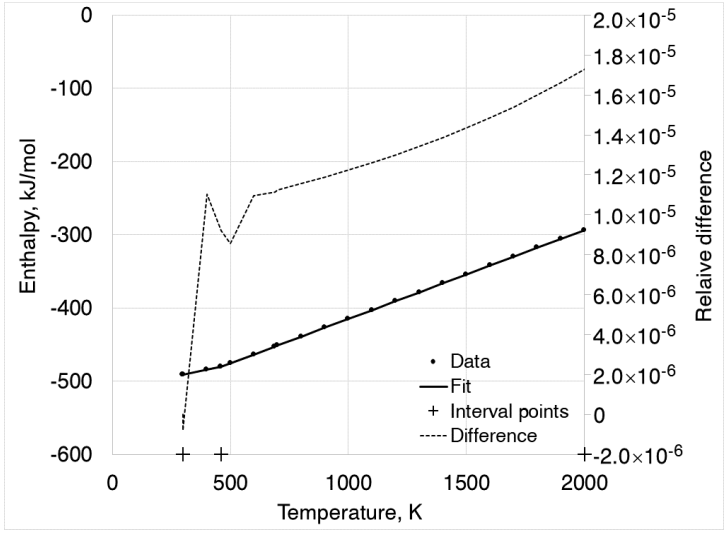


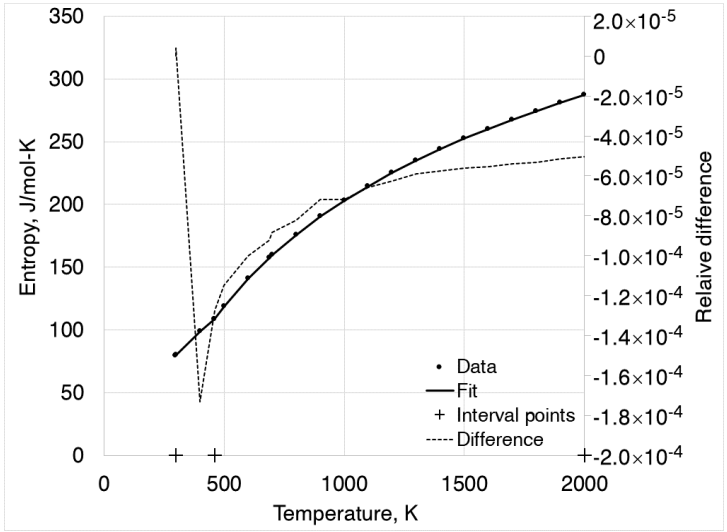


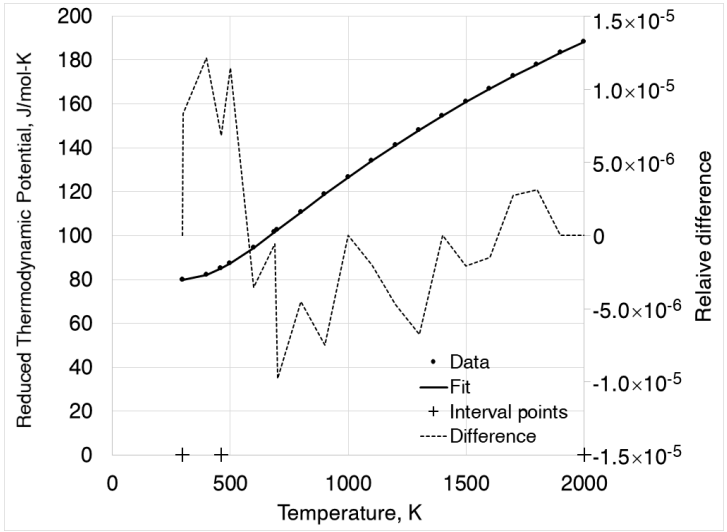
Beryllium Chloride ($BeCl_2$) – Be1Cl2(l)

MW = 79.918 g/mol – $\Delta_f H^o$(298.15 K) = -491.313 kJ/mol – S^o(298.15 K) = 79.727 J/mol-K – p^o = 0.1 MPa



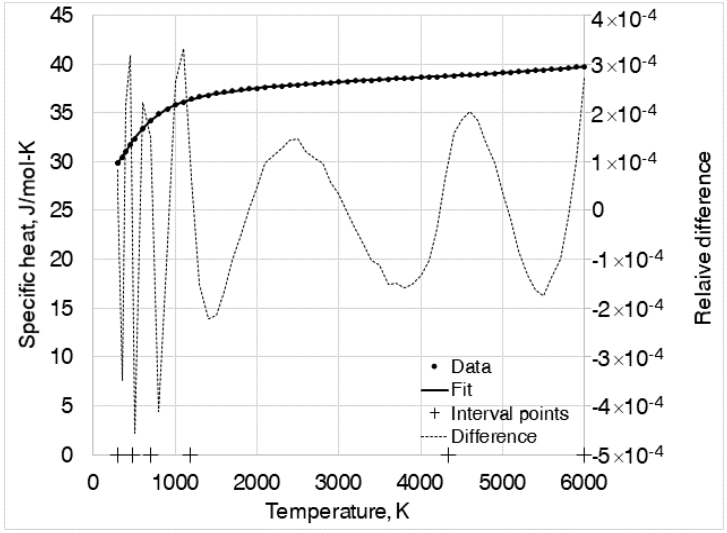


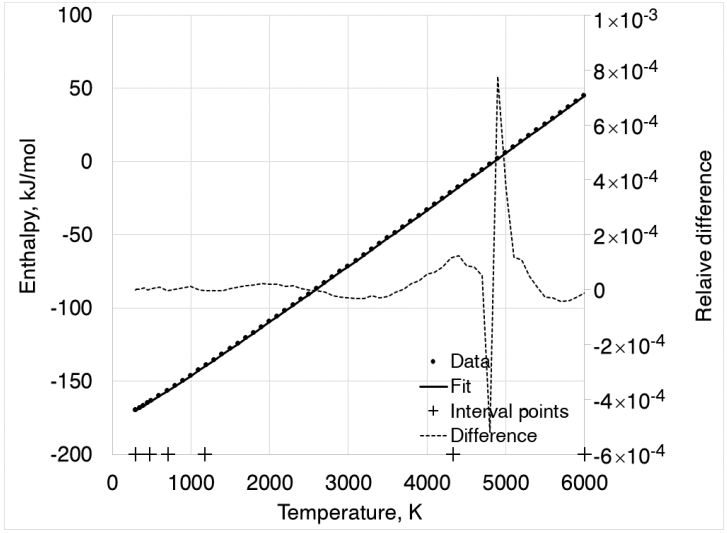


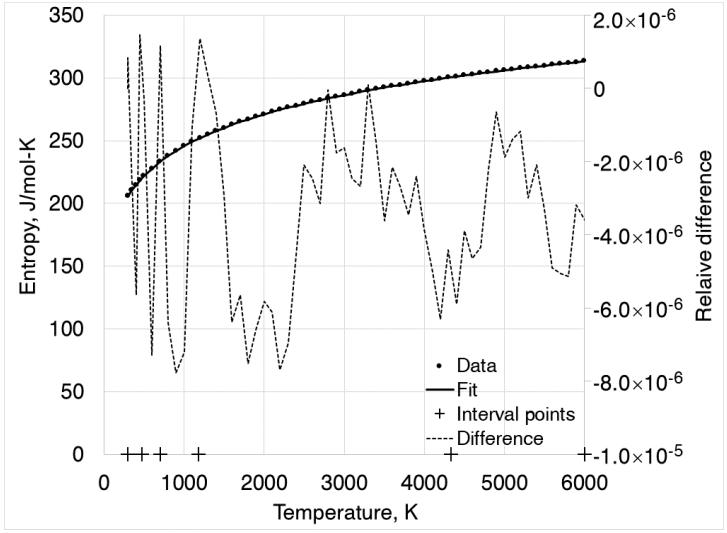


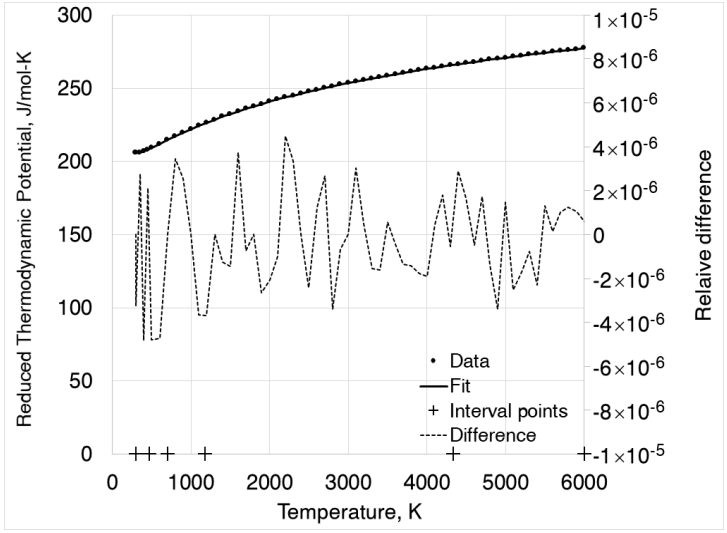
Beryllium Fluoride (BeF) – Be1F1(g)

MW = 28.0106 g/mol – $\Delta_f H^o$(298.15 K) = -169.87 kJ/mol – S^o(298.15 K) = 205.751 J/mol-K – p^o = 0.1 MPa



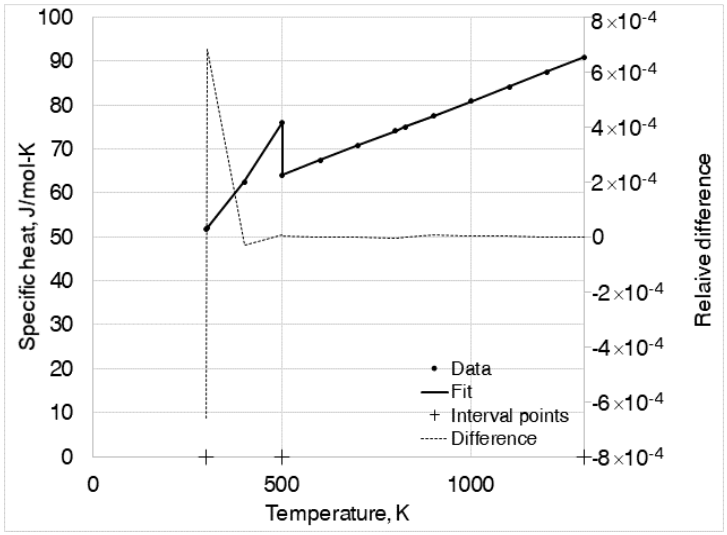


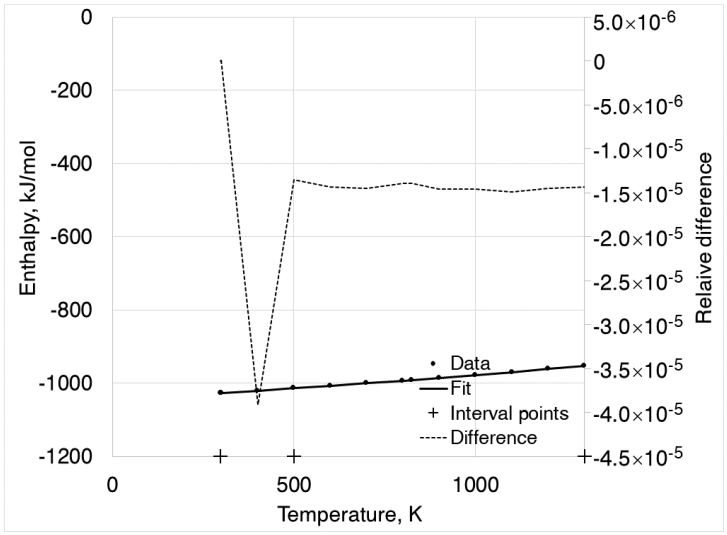


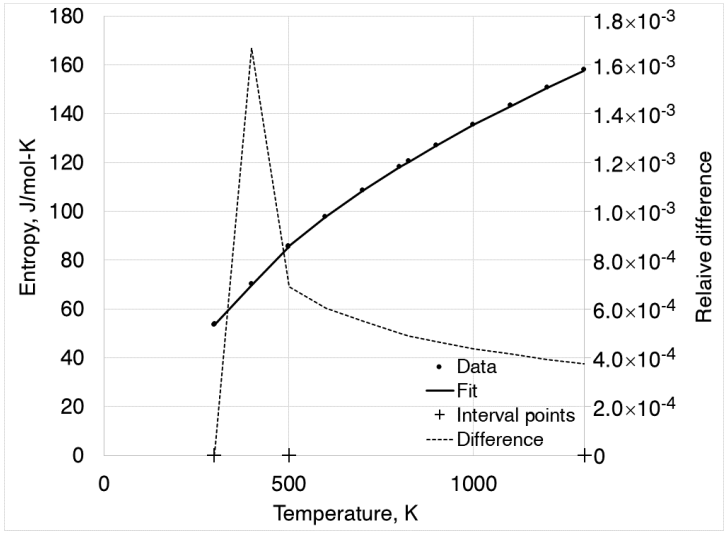


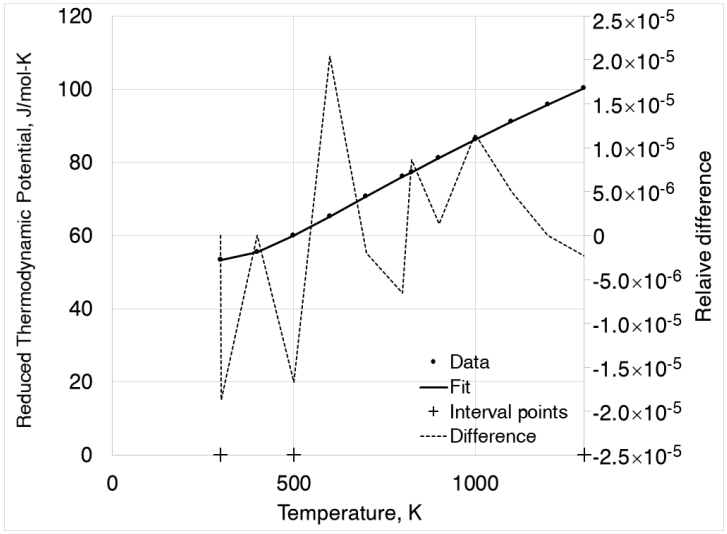
Beryllium Fluoride (BeF_2) – Be1F2(cr)

MW = 47.009 g/mol – $\Delta_f H^o$(298.15 K) = -1026.754 kJ/mol – S^o(298.15 K) = 53.354 J/mol-K – p^o = 0.1 MPa



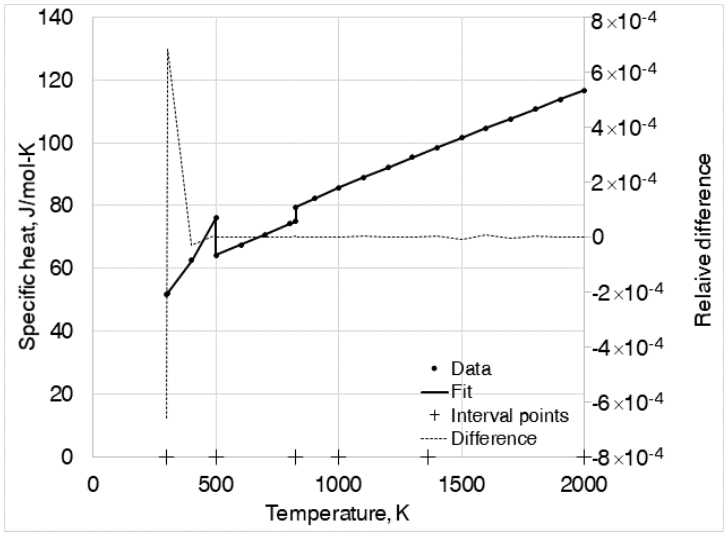


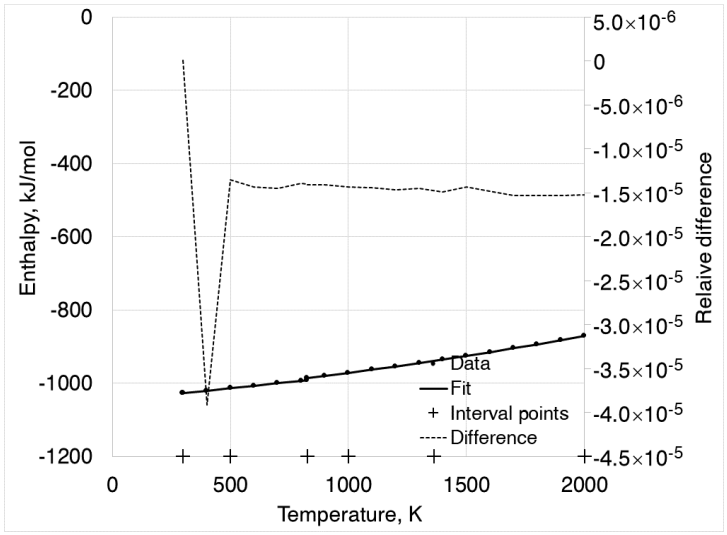


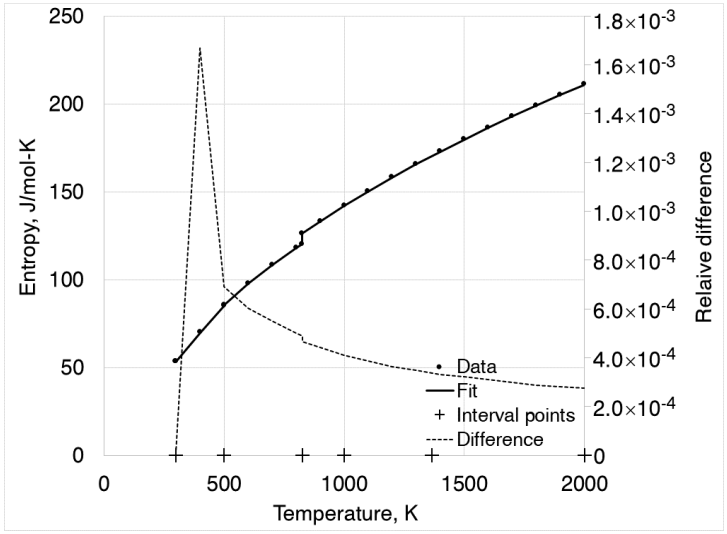


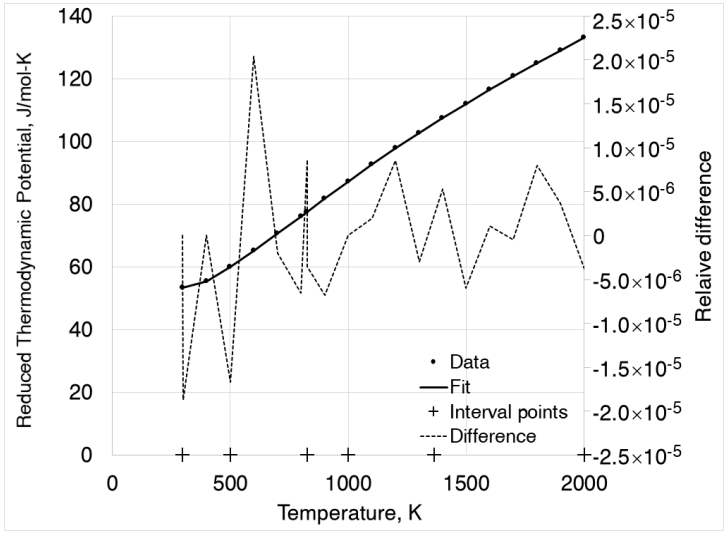
Beryllium Fluoride (BeF_2) – Be1F2(cr,l)

MW = 47.009 g/mol – $\Delta_f H^o$(298.15 K) = -1026.754 kJ/mol – S^o(298.15 K) = 53.354 J/mol-K – p^o = 0.1 MPa



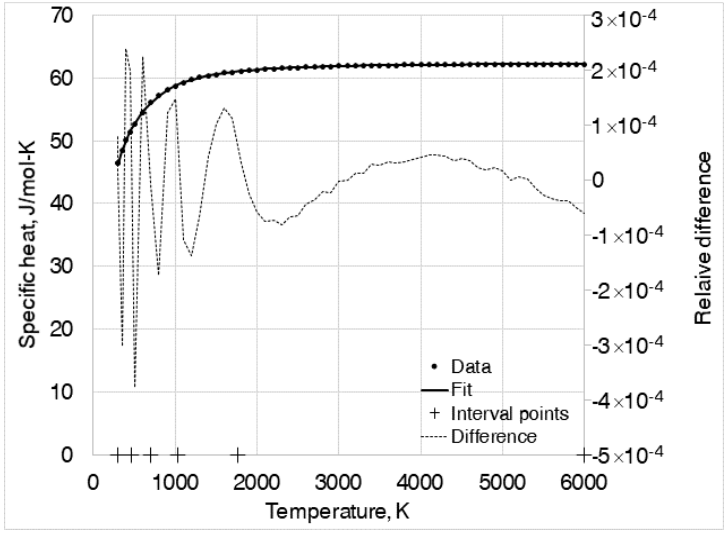


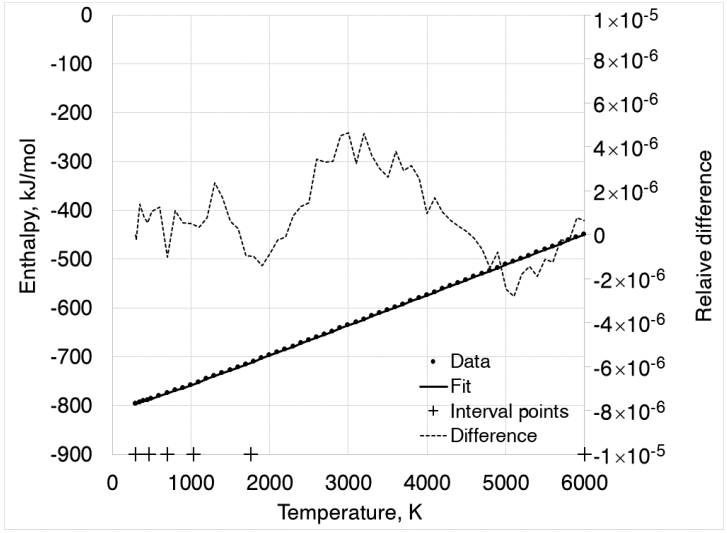


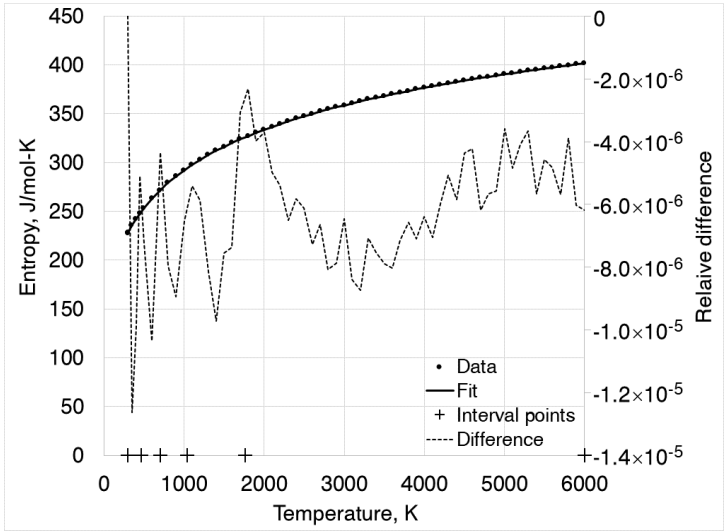


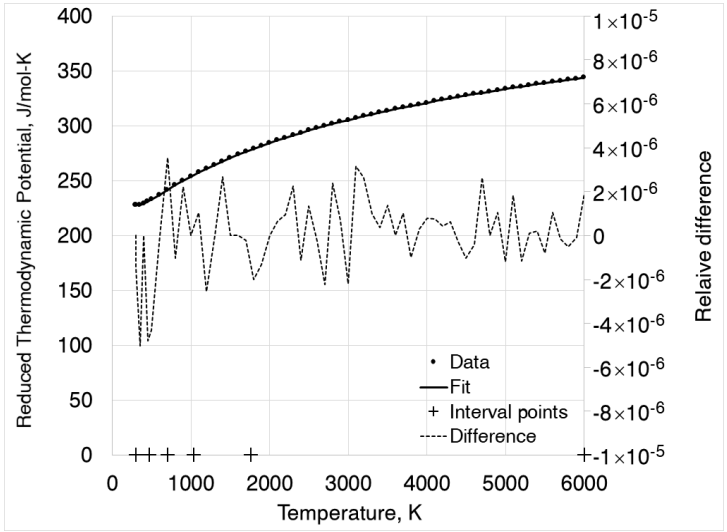
Beryllium Fluoride (BeF_2) – Be1F2(g)

MW = 47.009 g/mol – $\Delta_f H^o$(298.15 K) = -796.006 kJ/mol – S^o(298.15 K) = 227.556 J/mol-K – p^o = 0.1 MPa



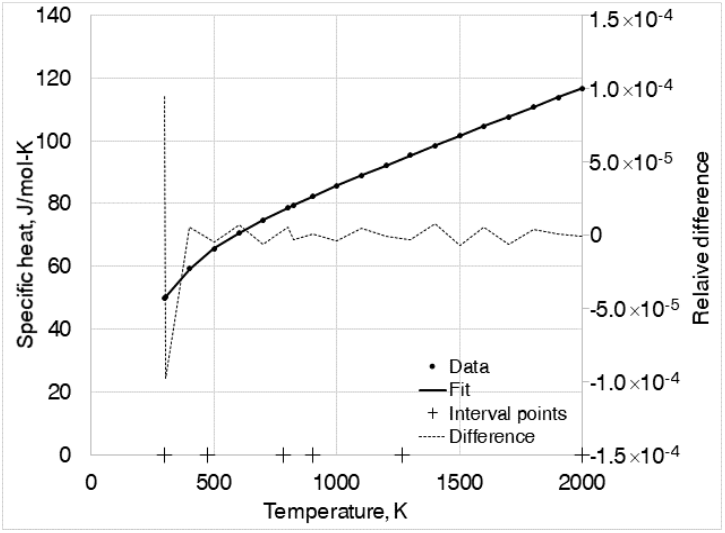


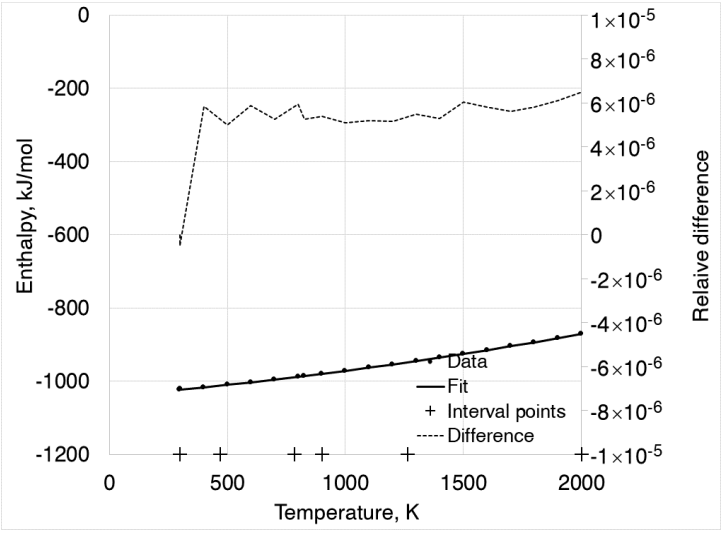


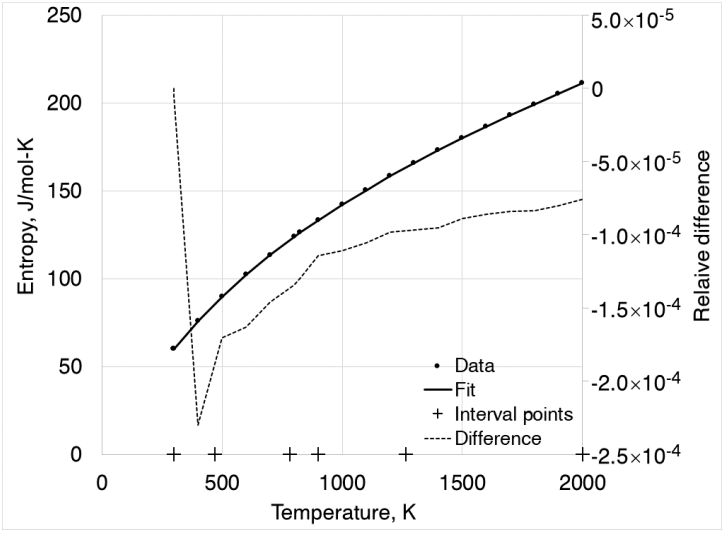


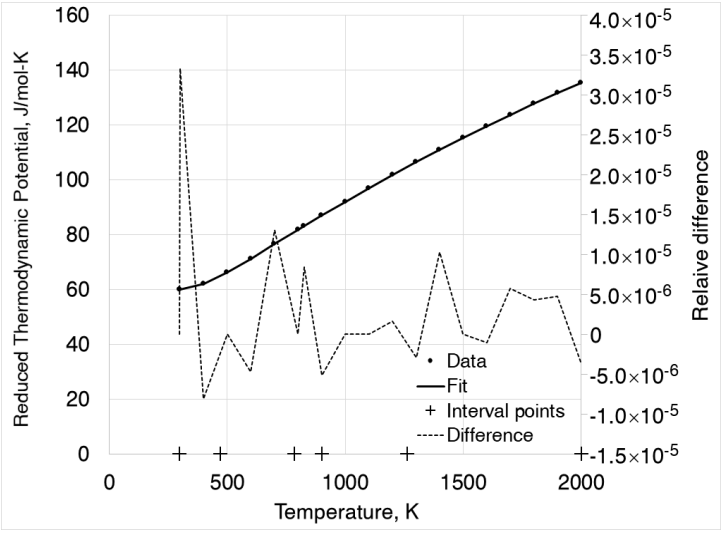
Beryllium Fluoride (BeF_2) – Be1F2(l)

MW = 47.009 g/mol – $\Delta_f H^o$(298.15 K) = -1022.048 kJ/mol – S^o(298.15 K) = 59.92 J/mol-K – p^o = 0.1 MPa



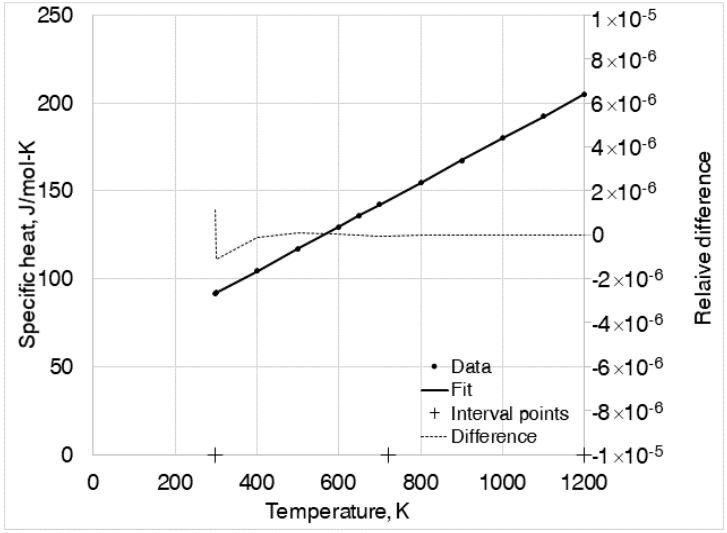


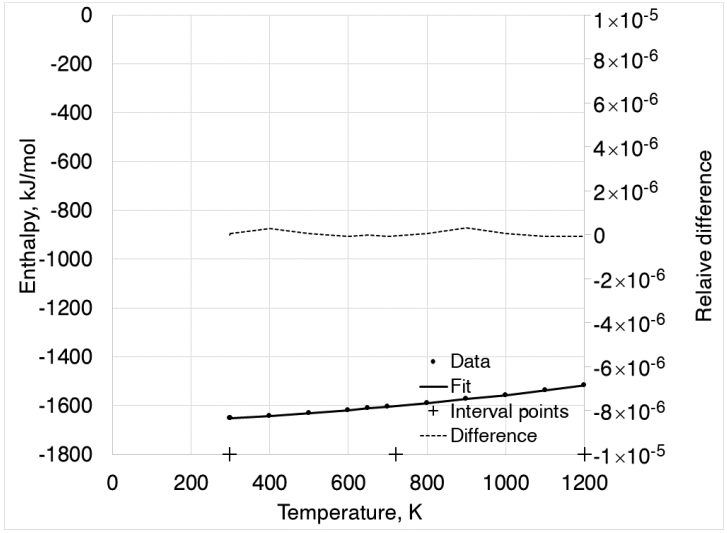


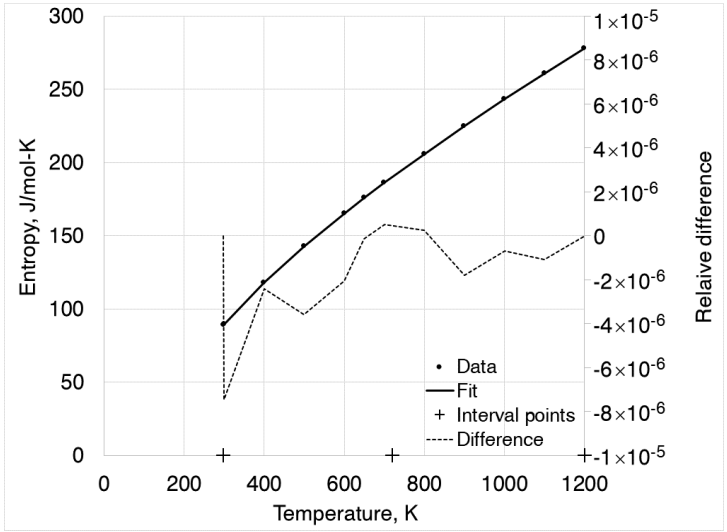


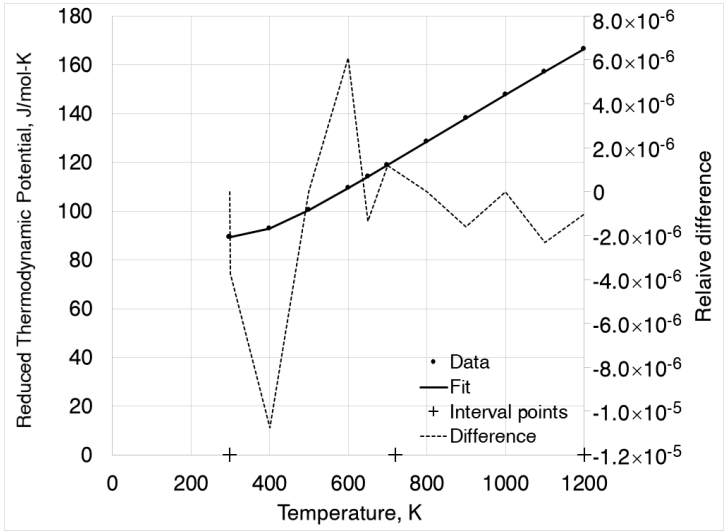
Lithium Trifluoroberyllate ($LiBeF_3$) – Be1F3Li1(cr)

MW = 72.948 g/mol – $\Delta_f H^o$(298.15 K) = -1651.843 kJ/mol – S^o(298.15 K) = 89.245 J/mol-K – p^o = 0.1 MPa



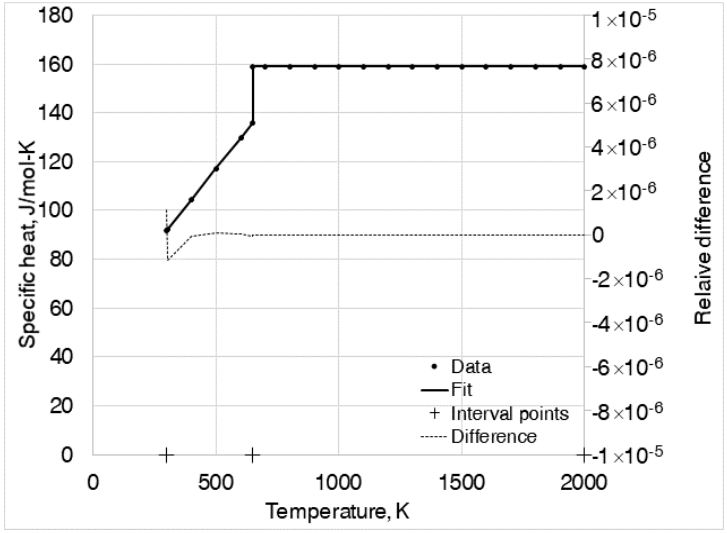


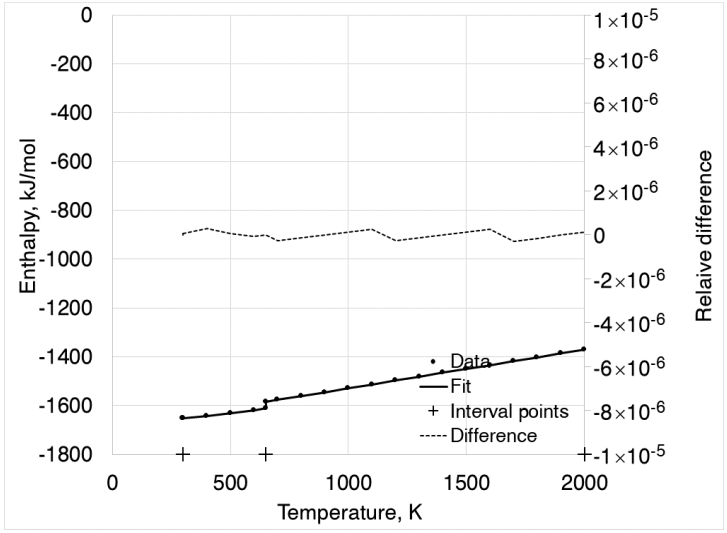


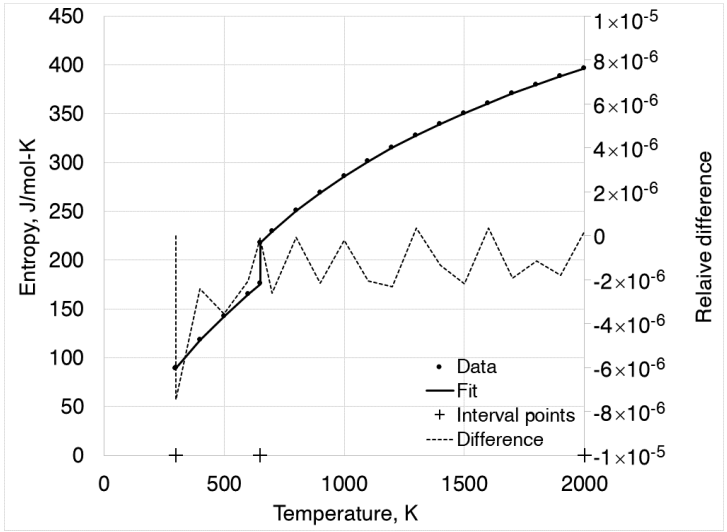


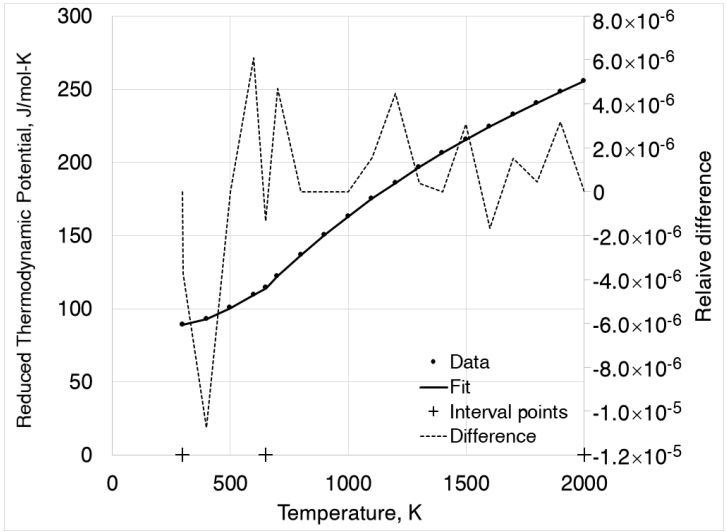
Lithium Trifluoroberyllate ($LiBeF_3$) – Be1F3Li1(cr,l)

MW = 72.948 g/mol – $\Delta_f H^o$(298.15 K) = -1651.843 kJ/mol – S^o(298.15 K) = 89.245 J/mol-K – p^o = 0.1 MPa



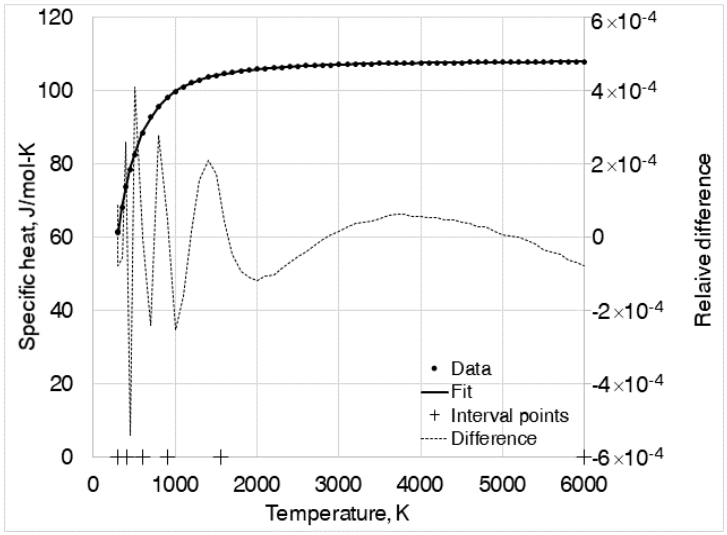


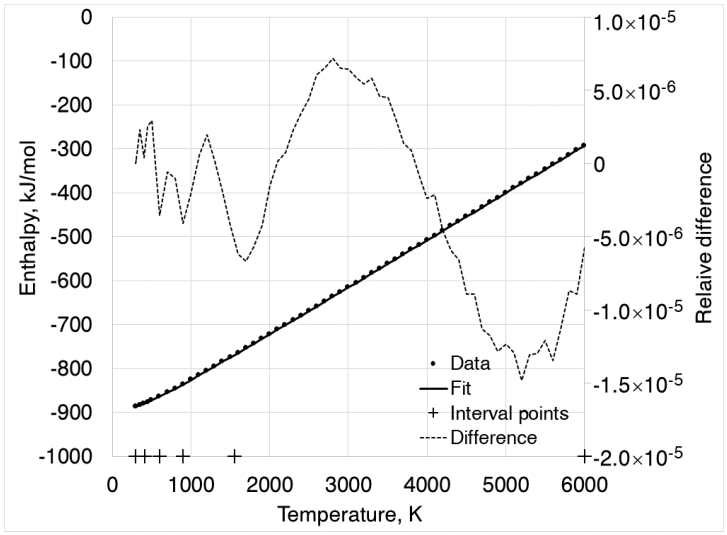


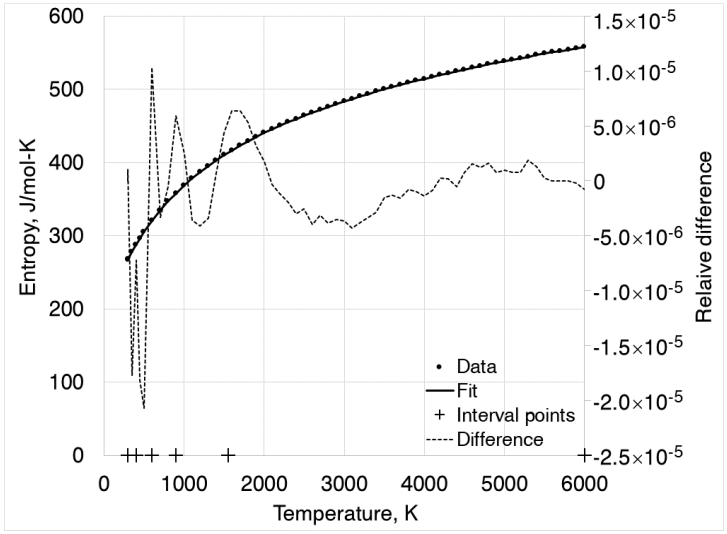


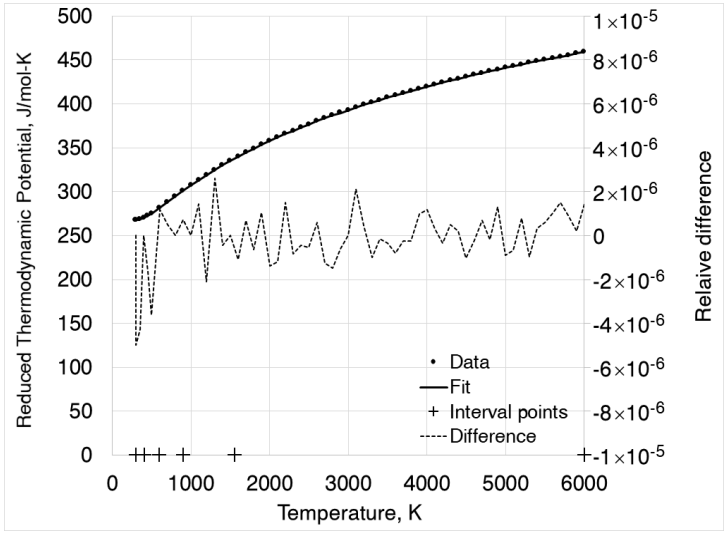
Lithium Trifluoroberyllate ($LiBeF_3$) – Be1F3Li1(g)

MW = 72.948 g/mol – $\Delta_f H^o$(298.15 K) = -887.008 kJ/mol – S^o(298.15 K) = 267.419 J/mol-K – p^o = 0.1 MPa



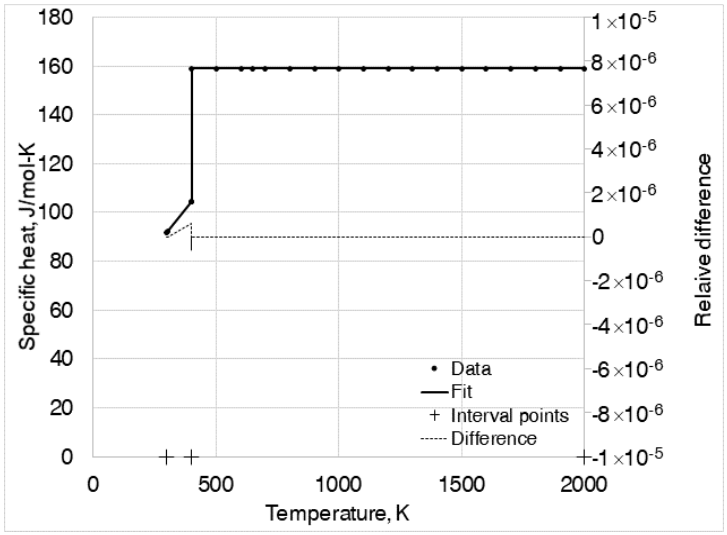


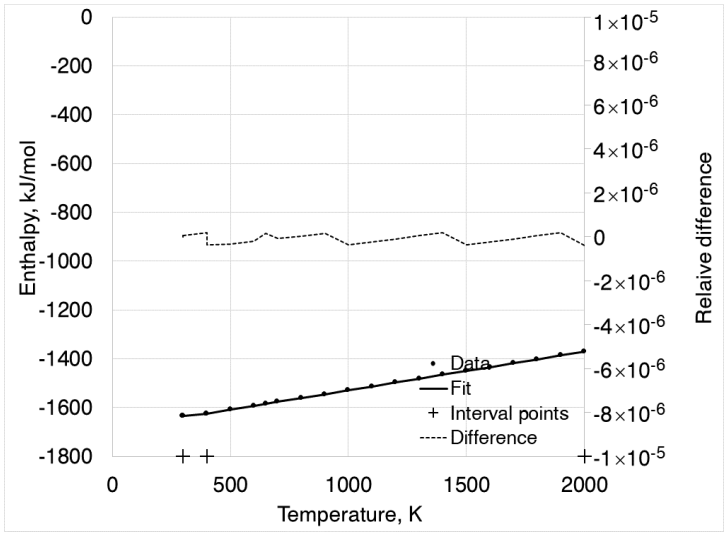


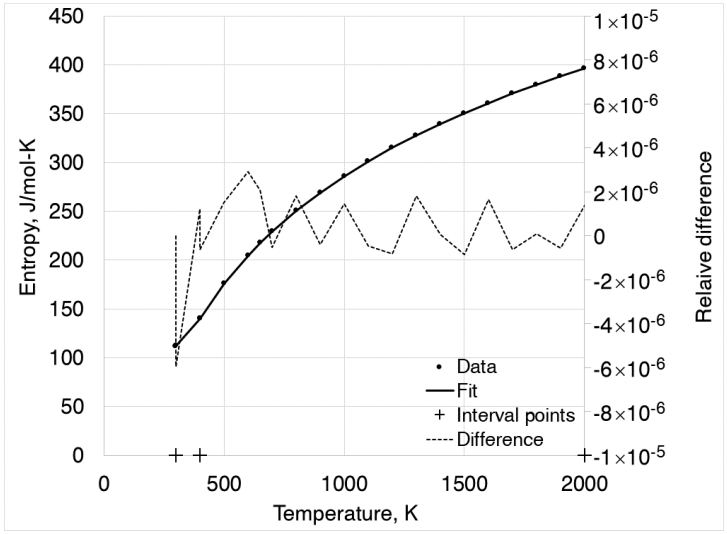


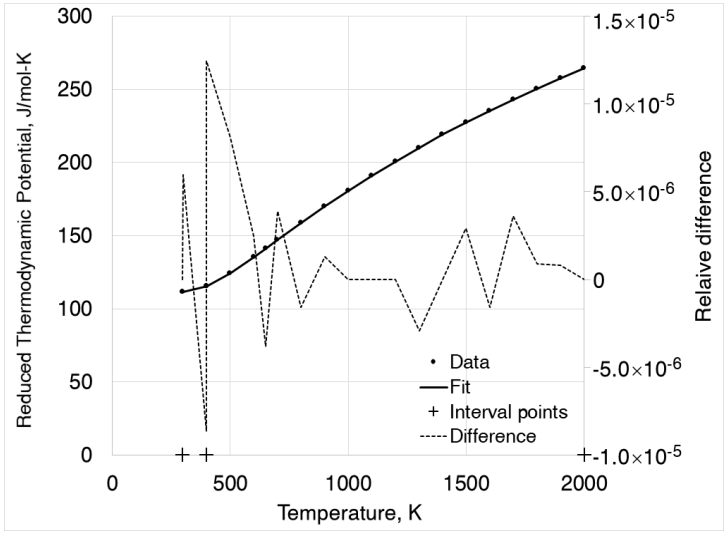
Lithium Trifluoroberyllate ($LiBeF_3$) – Be1F3Li1(l)

MW = 72.948 g/mol – $\Delta_f H^o$(298.15 K) = -1634.323 kJ/mol – S^o(298.15 K) = 111.681 J/mol-K – p^o = 0.1 MPa



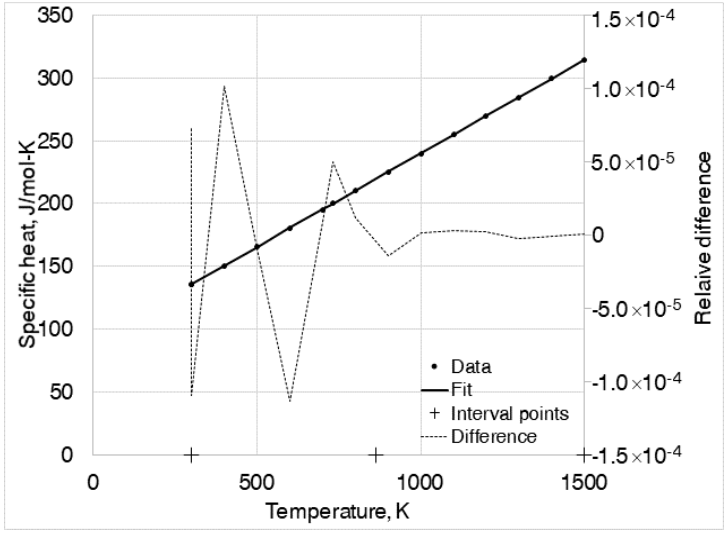


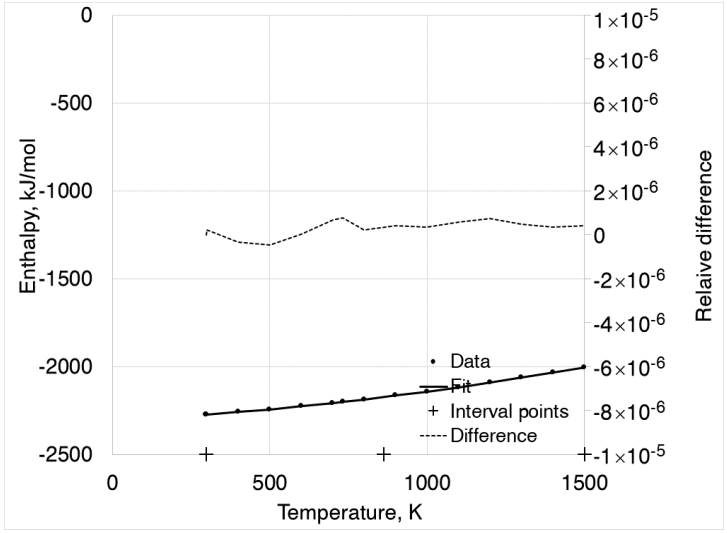


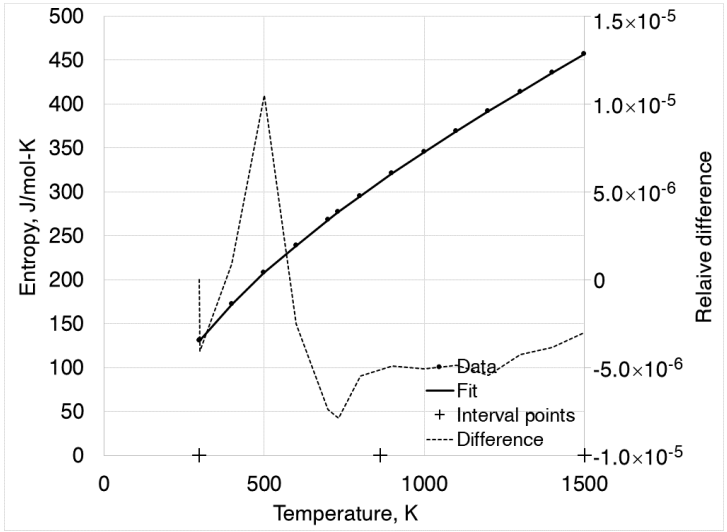


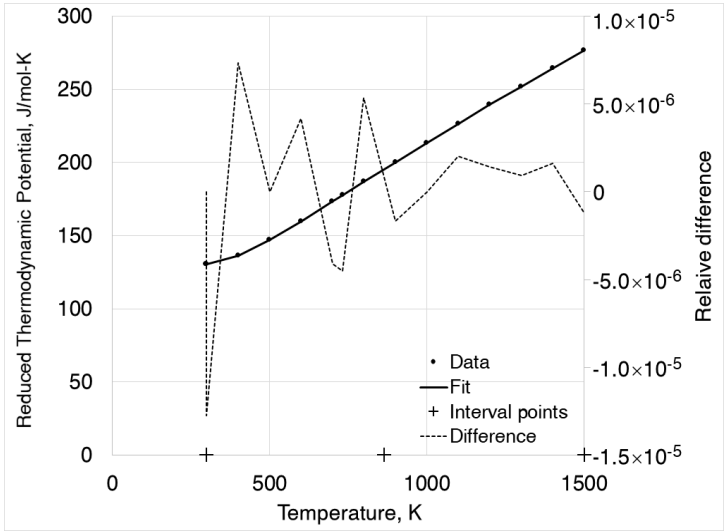
Lithium Tetrafluoroberyllate (Li_2BeF_4) – Be1F4Li2(cr)

MW = 98.888 g/mol – $\Delta_f H^o$(298.15 K) = -2273.586 kJ/mol – S^o(298.15 K) = 130.583 J/mol-K – p^o = 0.1 MPa



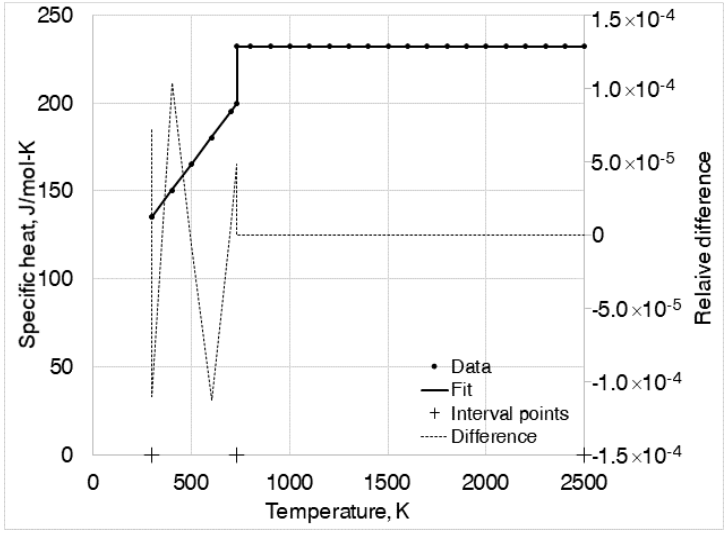


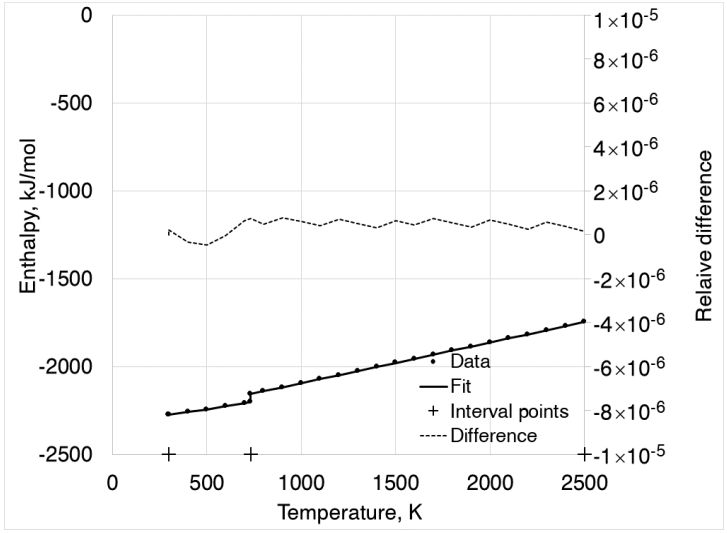


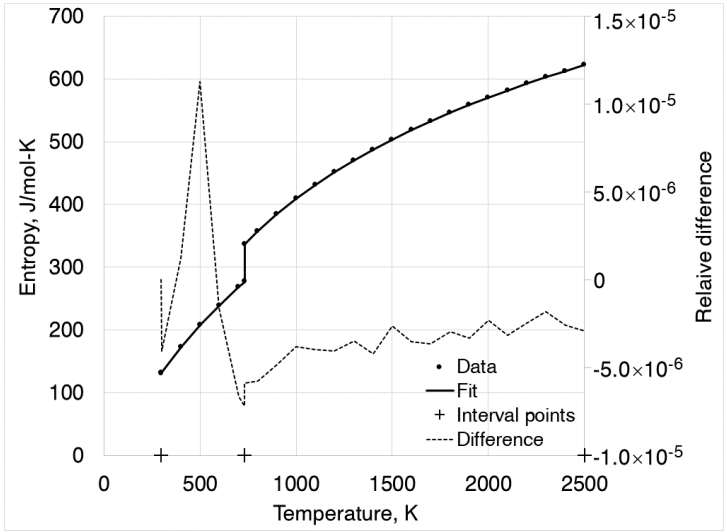


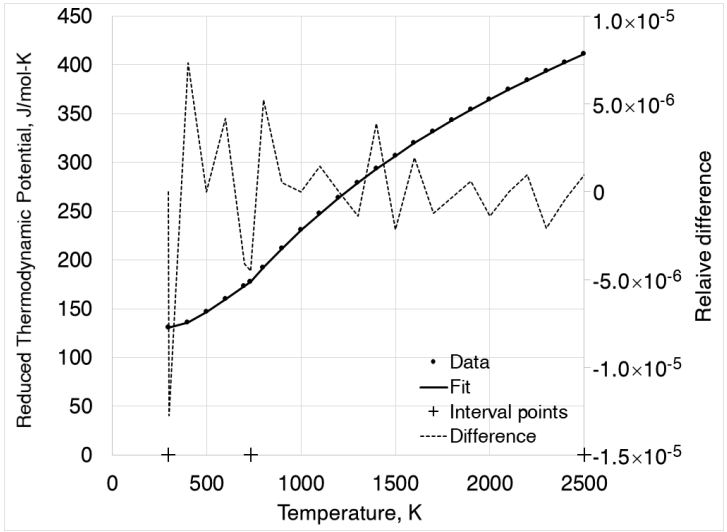
Lithium Tetrafluoroberyllate (Li_2BeF_4) – Be1F4Li2(cr,l)

MW = 98.888 g/mol – $\Delta_f H^o$(298.15 K) = -2273.586 kJ/mol – S^o(298.15 K) = 130.583 J/mol-K – p^o = 0.1 MPa



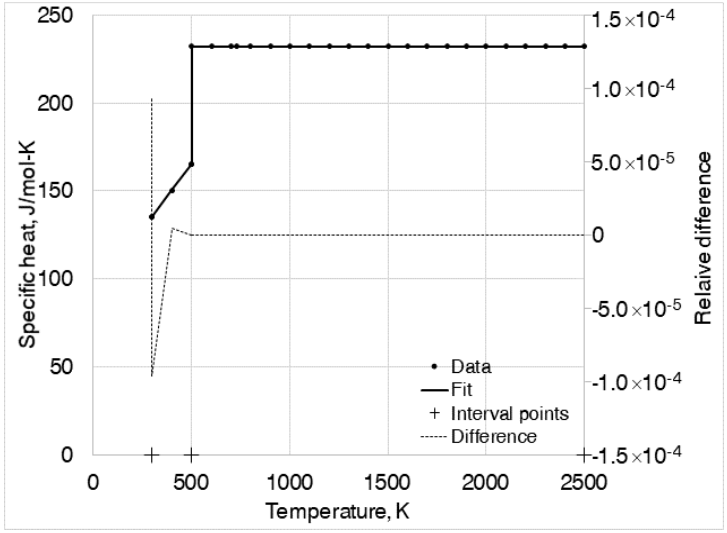


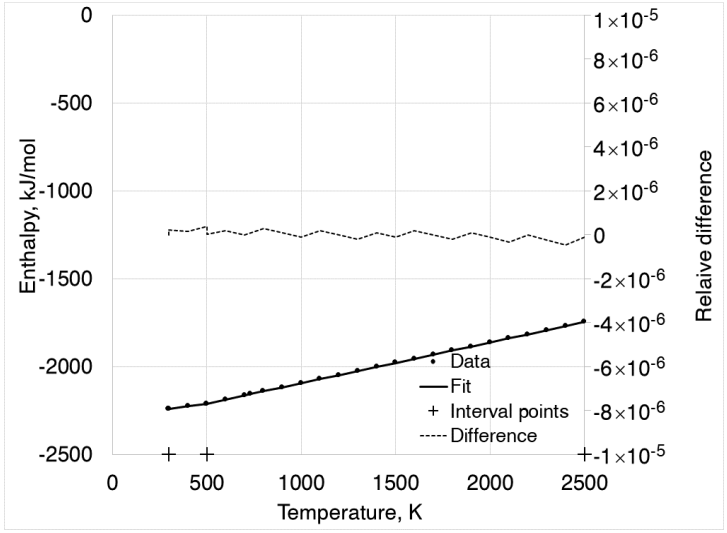


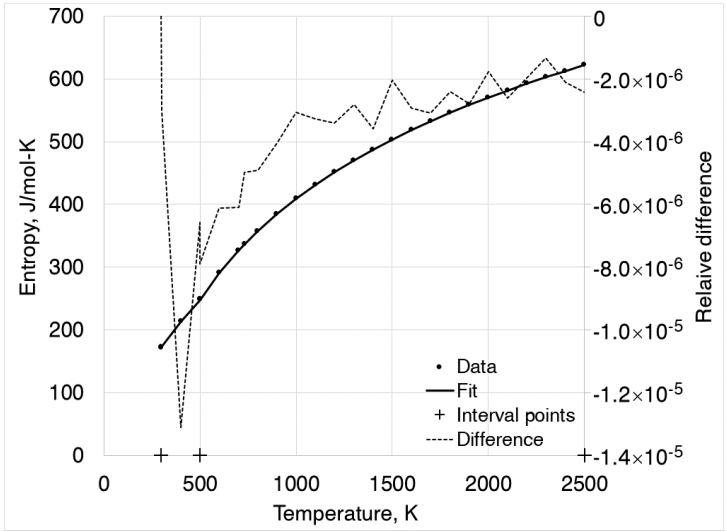


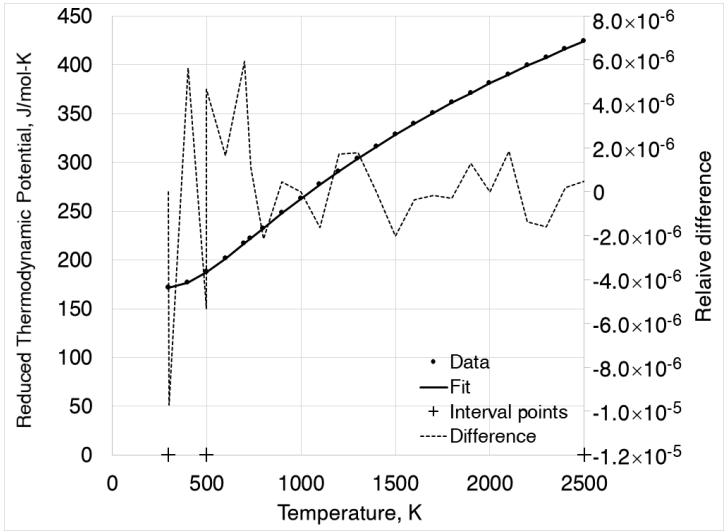
Lithium Tetrafluoroberyllate (Li_2BeF_4) – Be1F4Li2(l)

MW = 98.888 g/mol – $\Delta_f H^o$(298.15 K) = -2241.068 kJ/mol – S^o(298.15 K) = 171.395 J/mol-K – p^o = 0.1 MPa



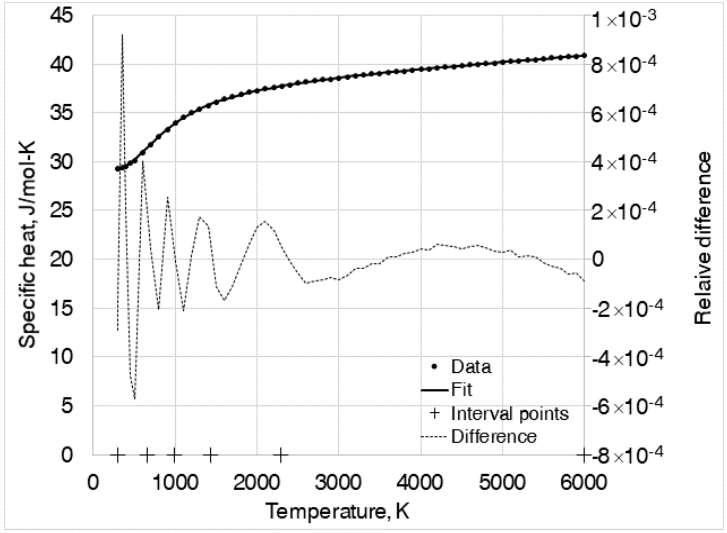


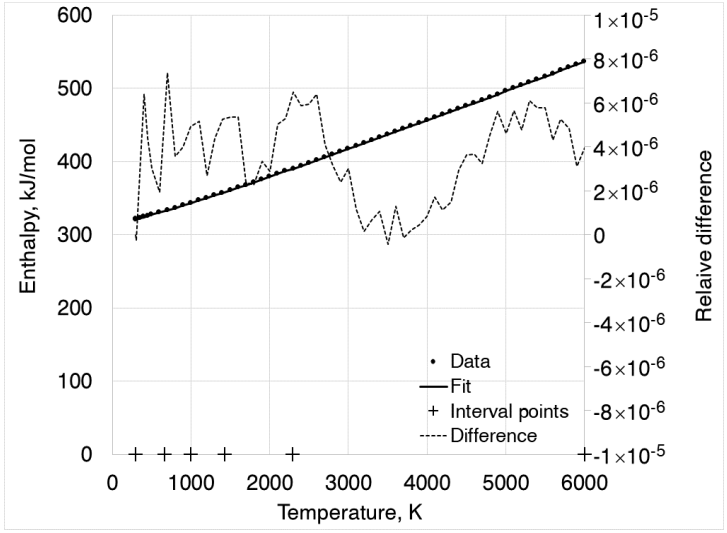


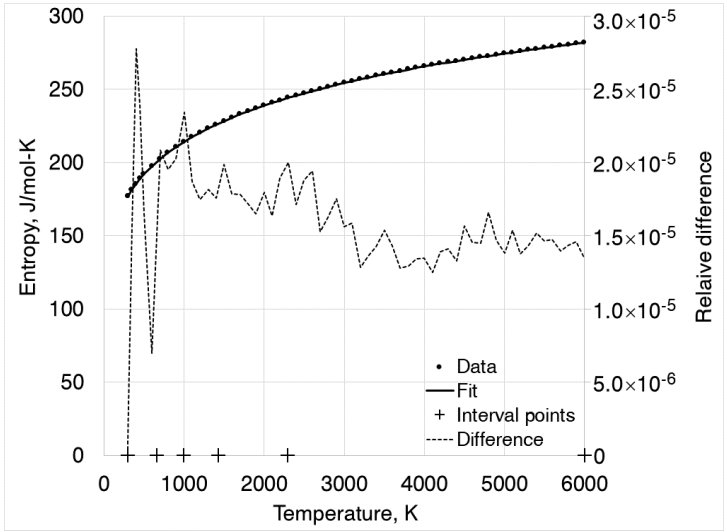


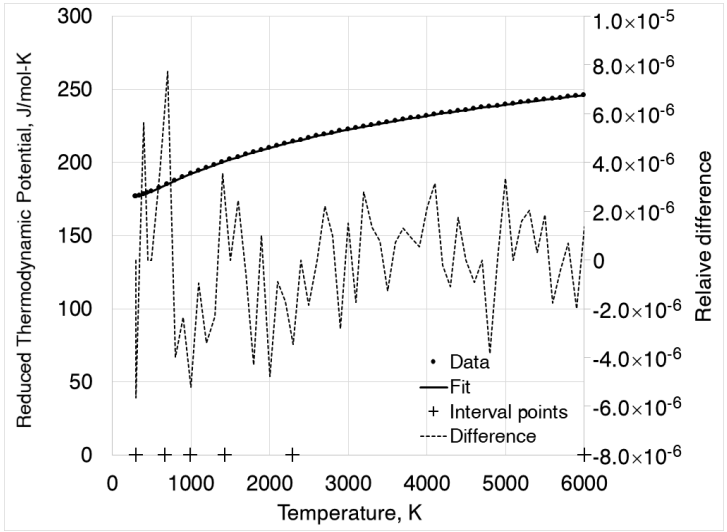
Beryllium Hydride (BeH) – Be1H1(g)

MW = 10.02012 g/mol – $\Delta_f H^o$(298.15 K) = 321.197 kJ/mol – S^o(298.15 K) = 176.822 J/mol-K – p^o = 0.1 MPa



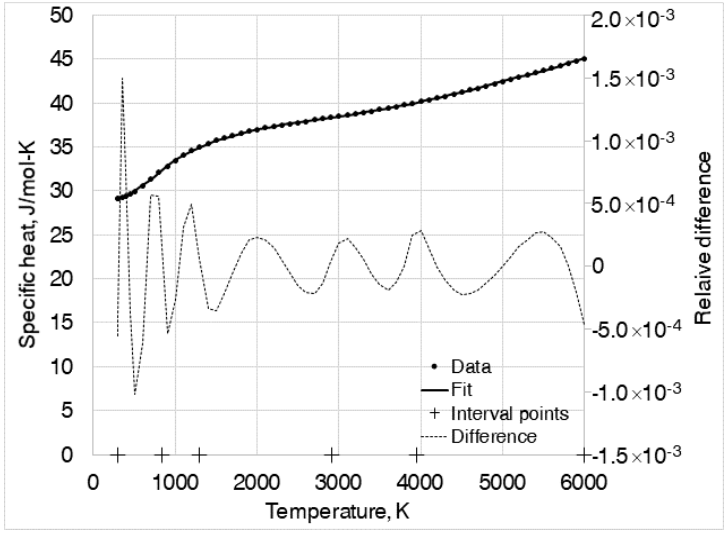


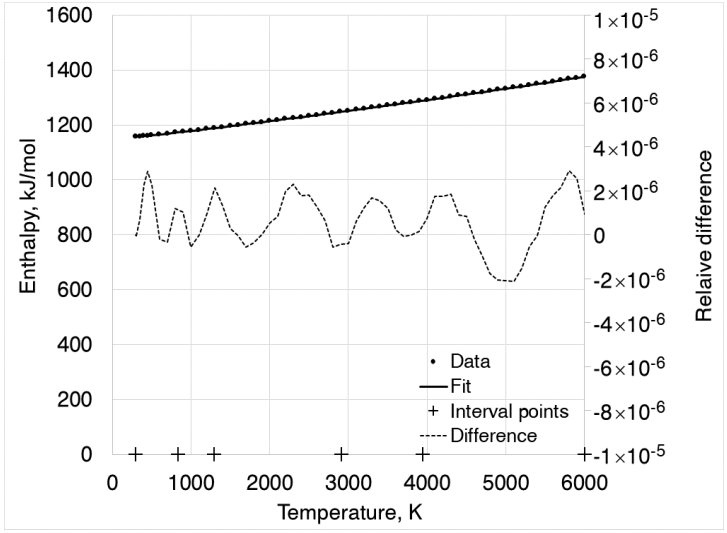


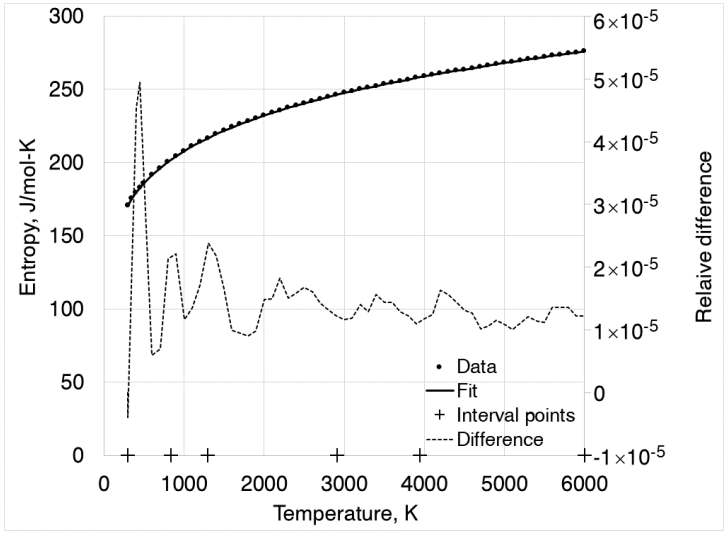


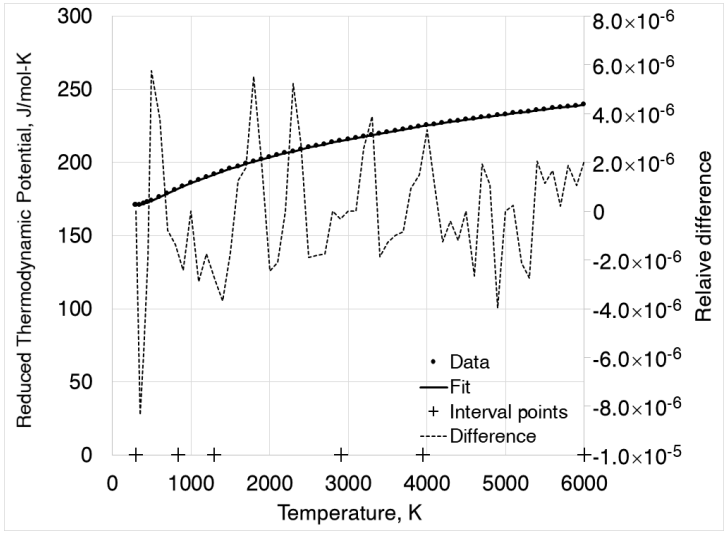
Beryllium Hydride, Ion (BeH^+) – Be1H1+(g)

MW = 10.01957 g/mol – $\Delta_f H^o$(298.15 K) = 1156.458 kJ/mol – S^o(298.15 K) = 170.653 J/mol-K – p^o = 0.1 MPa



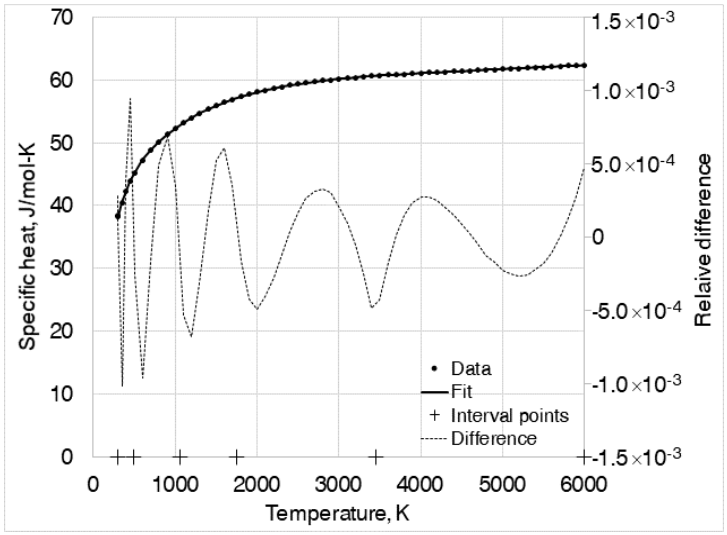


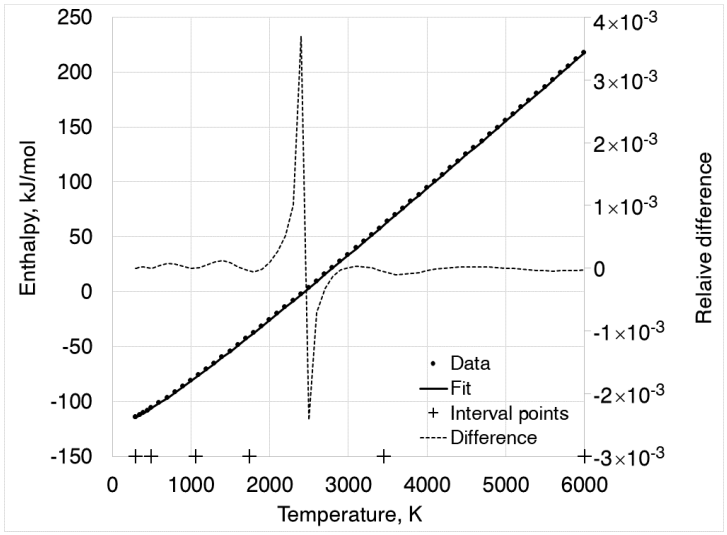


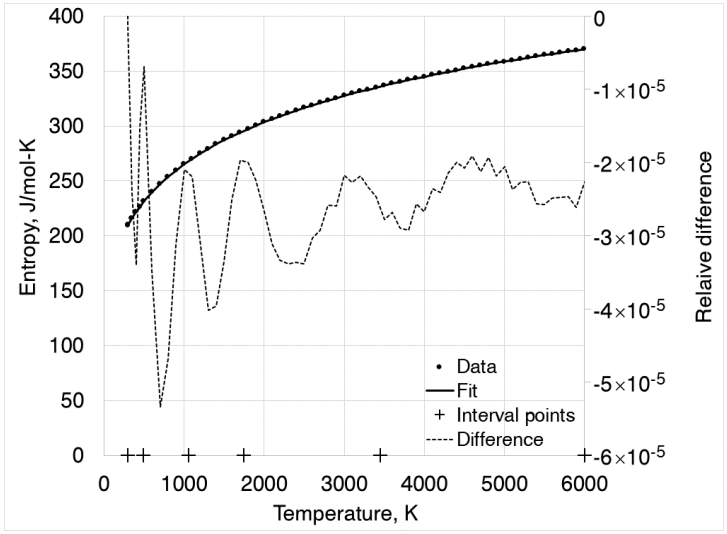


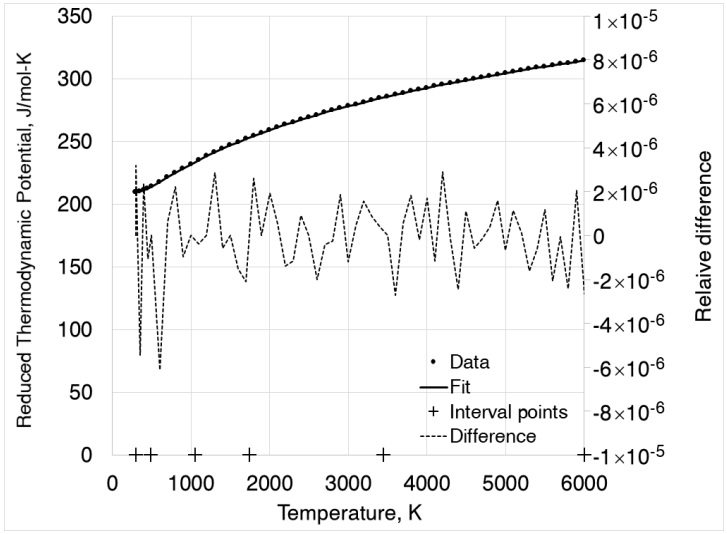
Beryllium Hydroxide (BeOH) – Be1H1O1(g)

MW = 26.0195 g/mol – $\Delta_f H^o$(298.15 K) = -114.642 kJ/mol – S^o(298.15 K) = 209.594 J/mol-K – p^o = 0.1 MPa



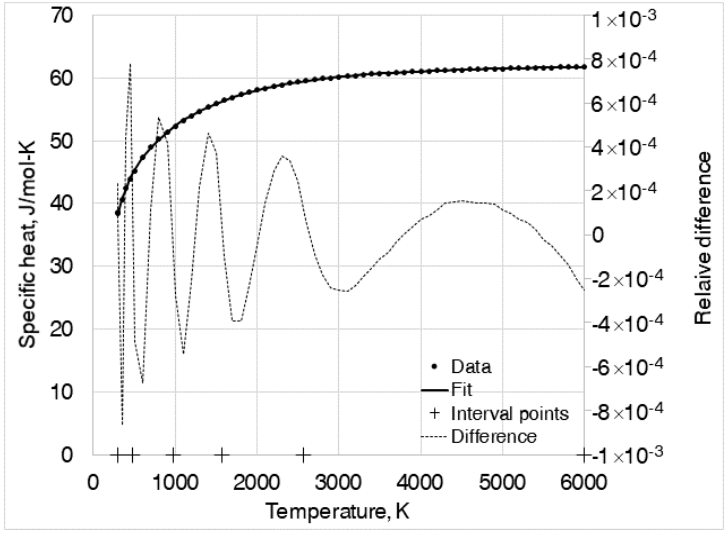


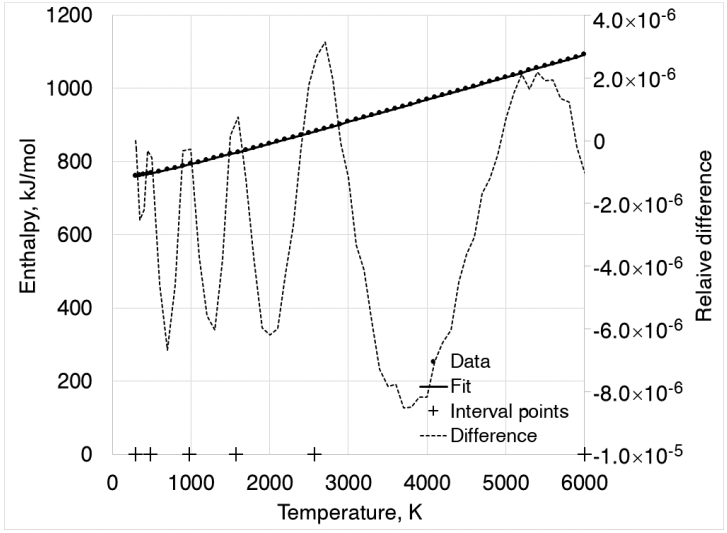


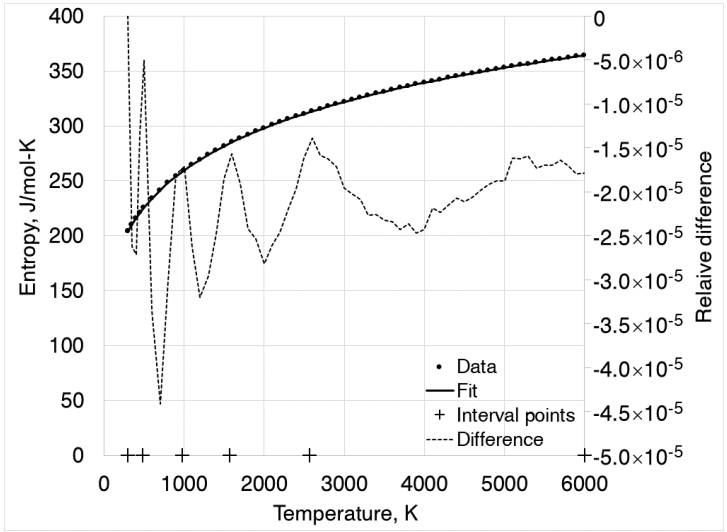


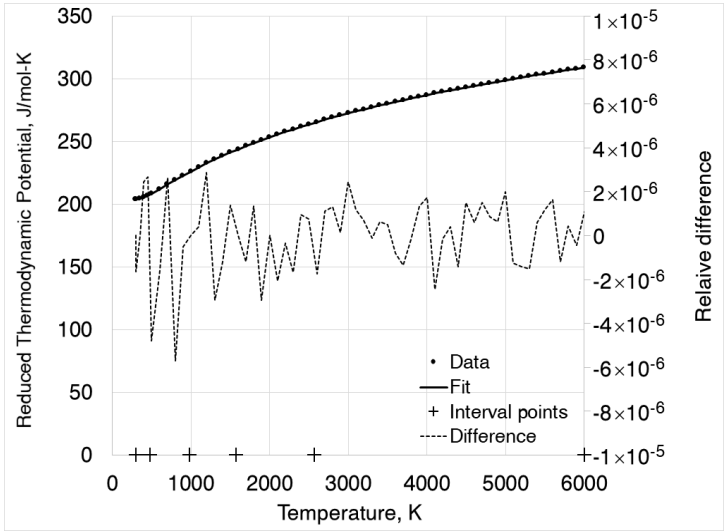
Beryllium Hydroxide, Ion ($BeOH^+$) – Be1H1O1+(g)

MW = 26.019 g/mol – $\Delta_f H^o$(298.15 K) = 759.934 kJ/mol – S^o(298.15 K) = 203.919 J/mol-K – p^o = 0.1 MPa



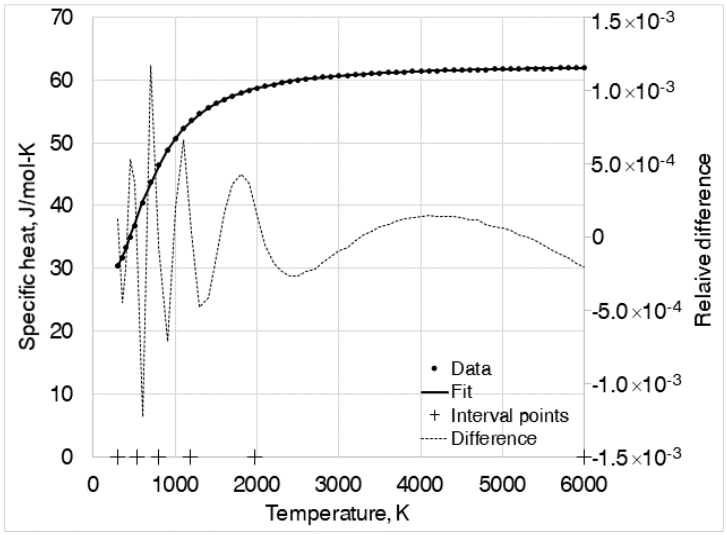


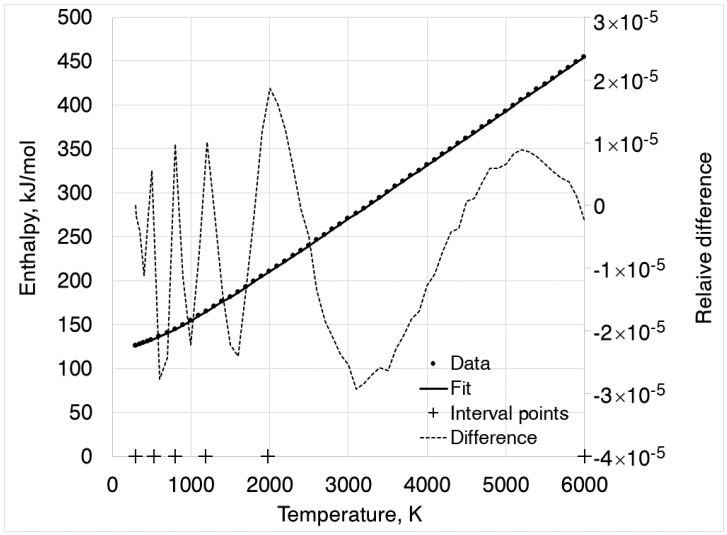


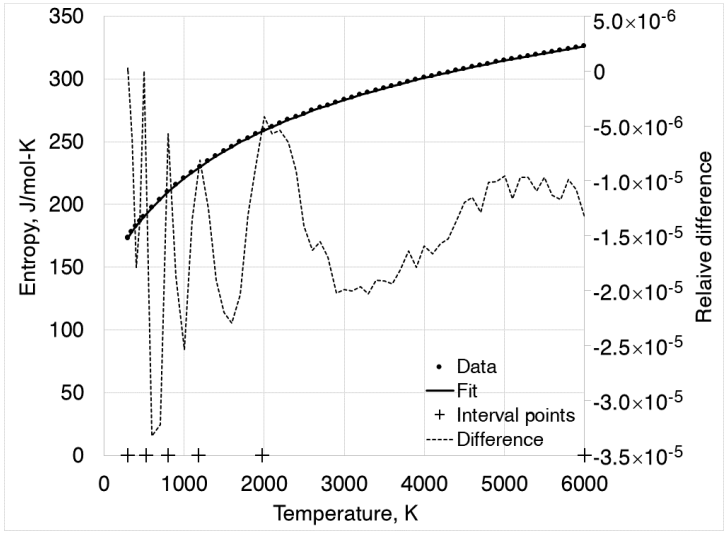


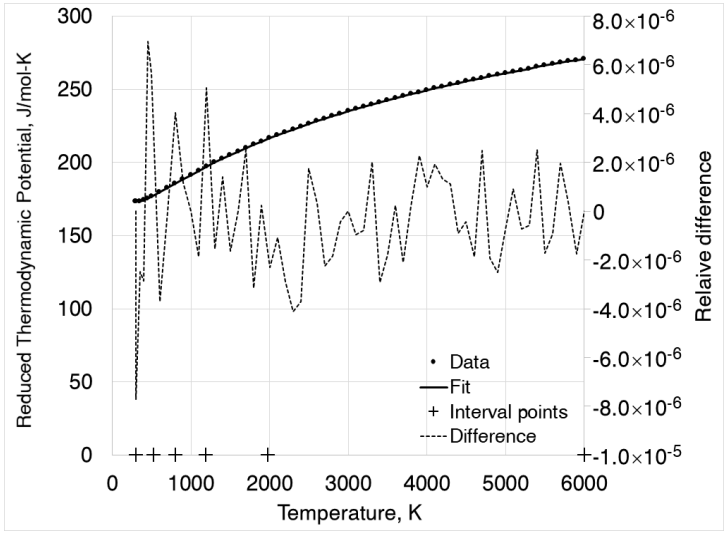
Beryllium Hydride (BeH_2) – Be1H2(g)

MW = 11.02806 g/mol – $\Delta_f H^o$(298.15 K) = 125.52 kJ/mol – S^o(298.15 K) = 173.115 J/mol-K – p^o = 0.1 MPa



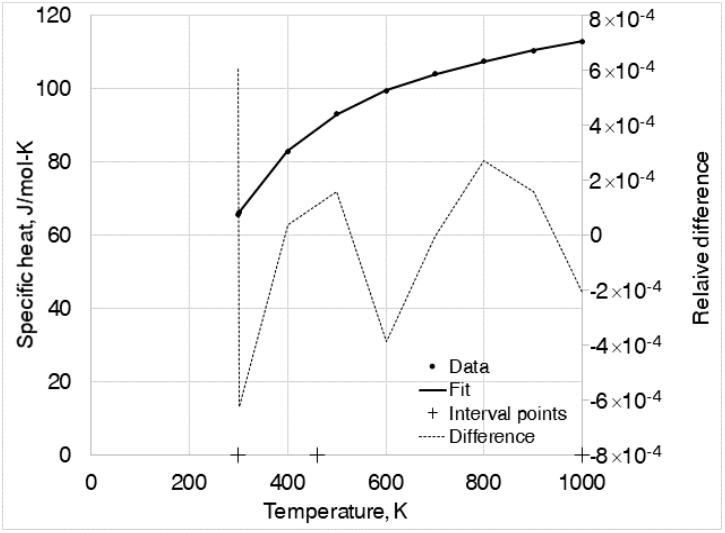


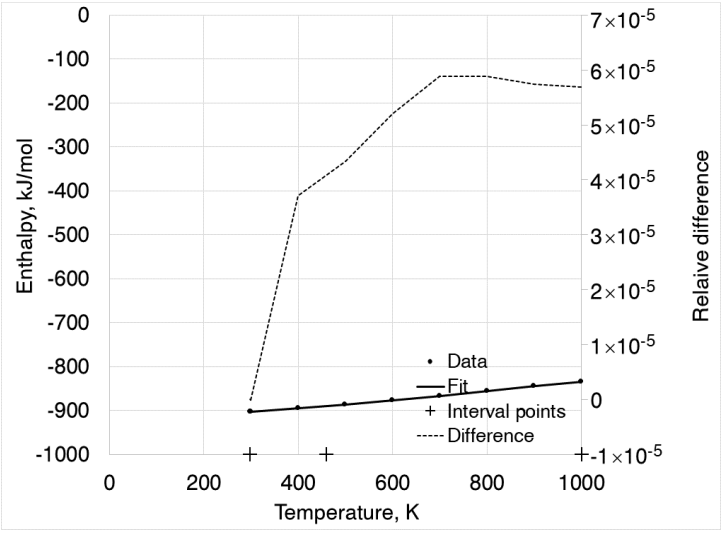


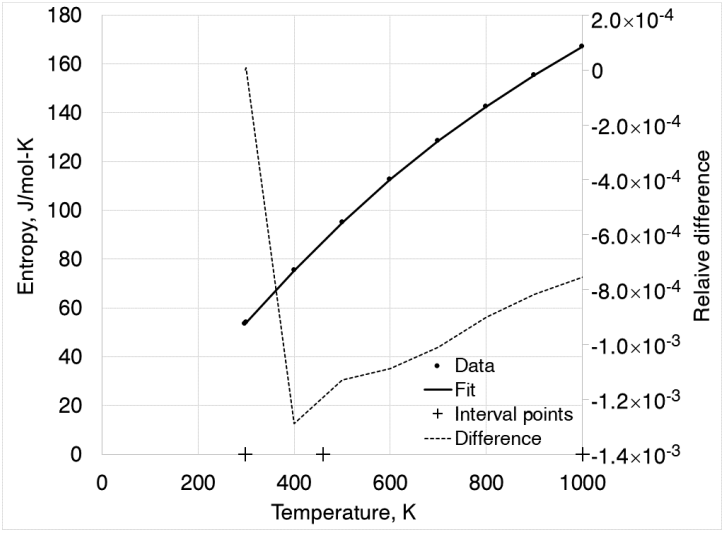


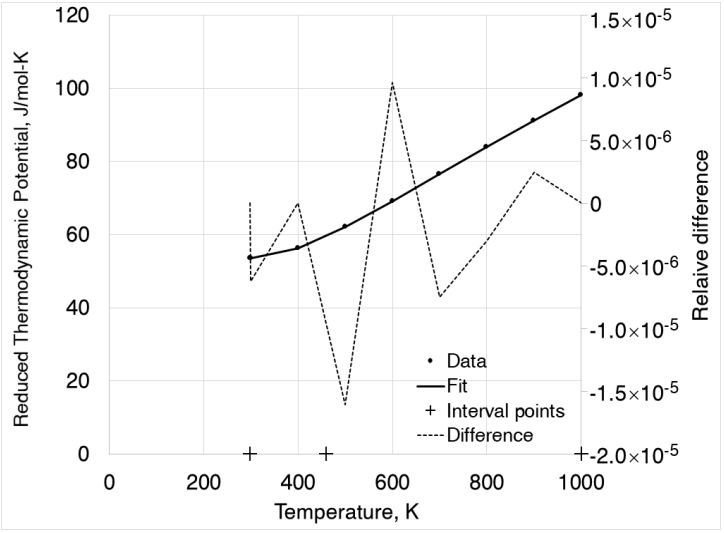
Beryllium Hydroxide, Alpha ($Be(OH)_2$) – Be1H2O2(cr)_01

MW = 43.0269 g/mol – $\Delta_f H^o$(298.15 K) = -902.907 kJ/mol – S^o(298.15 K) = 53.555 J/mol-K – p^o = 0.1 MPa



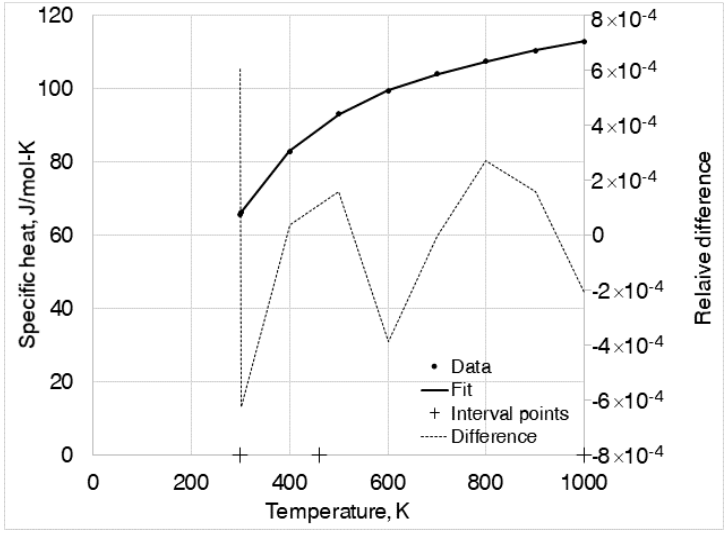


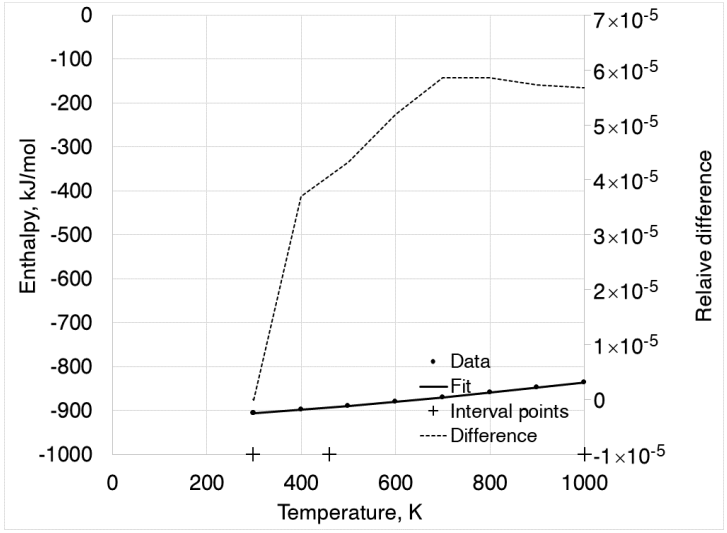


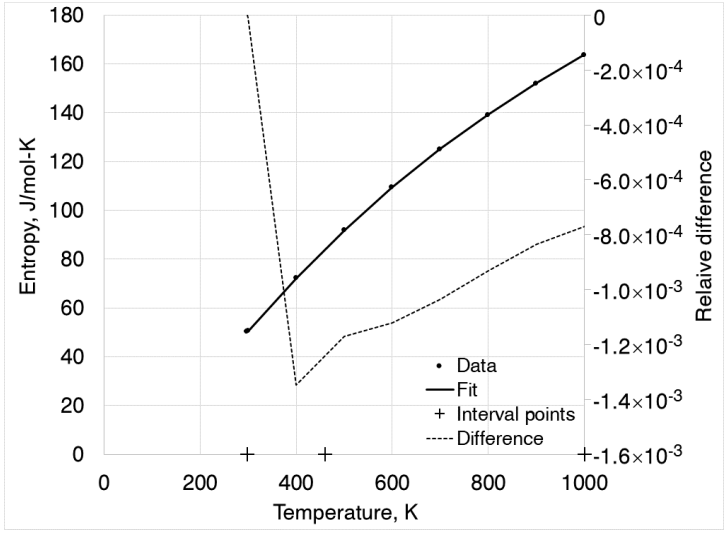


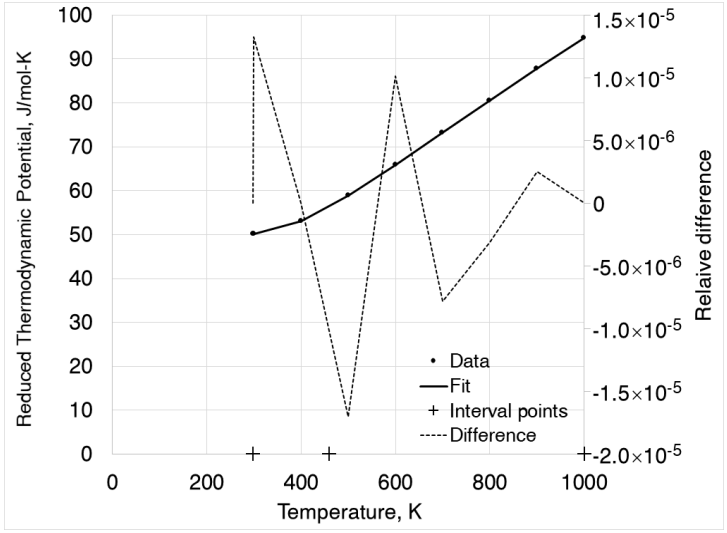
Beryllium Hydroxide, Beta ($Be(OH)_2$) – Be1H2O2(cr)_02

MW = 43.0269 g/mol – $\Delta_f H^o$(298.15 K) = -905.836 kJ/mol – S^o(298.15 K) = 50.208 J/mol-K – p^o = 0.1 MPa



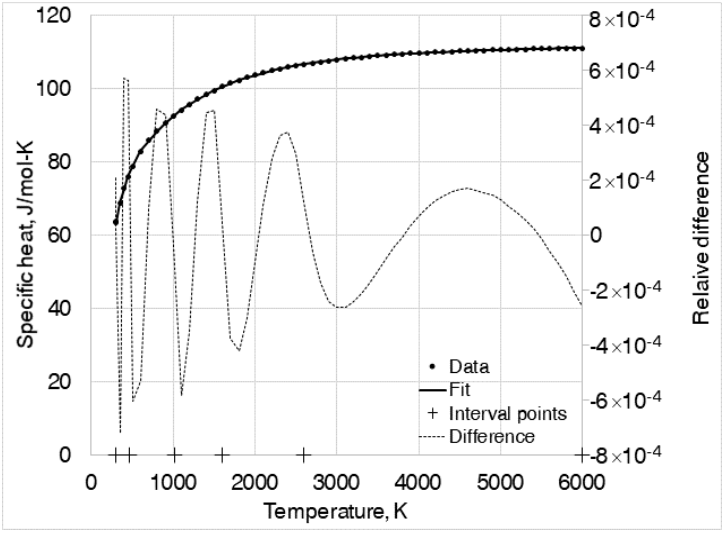


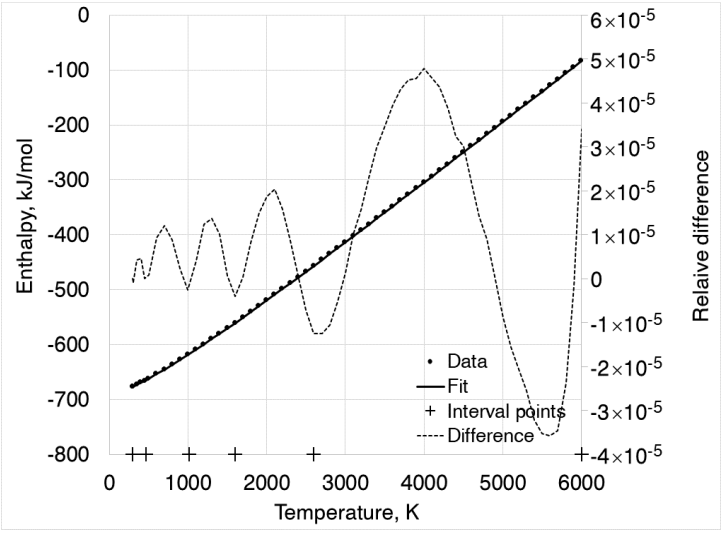


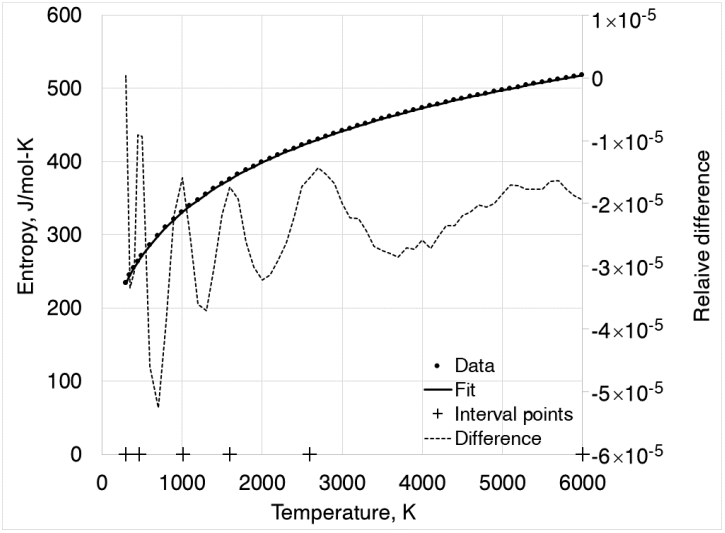


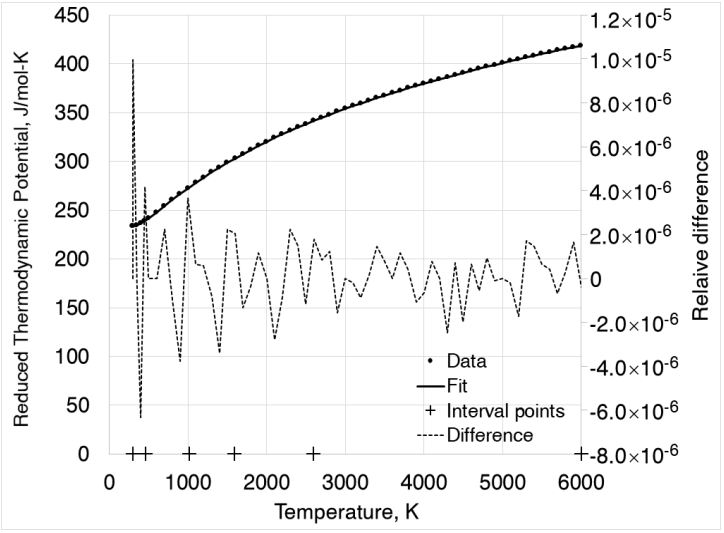
Beryllium Hydroxide ($Be(OH)_2$) – Be1H2O2(g)

MW = 43.0269 g/mol – $\Delta_f H^o$(298.15 K) = -676.553 kJ/mol – S^o(298.15 K) = 233.968 J/mol-K – p^o = 0.1 MPa



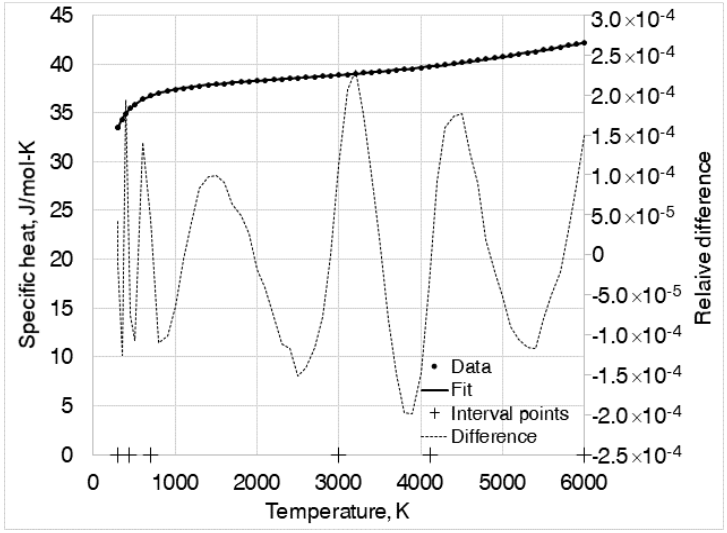


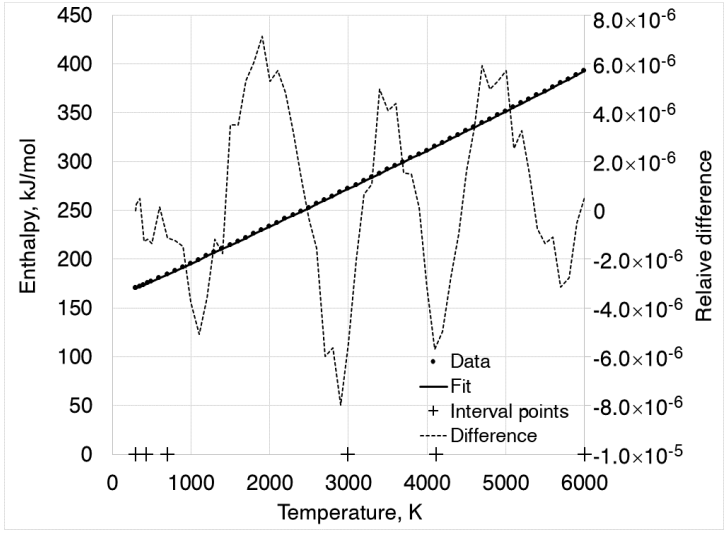


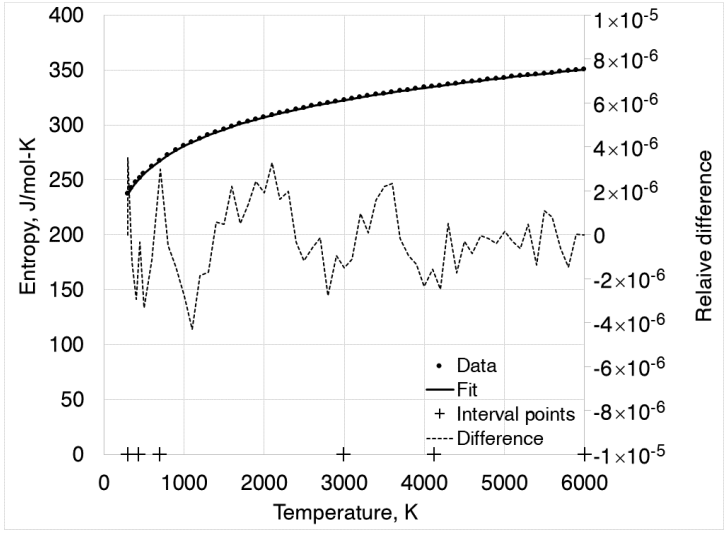


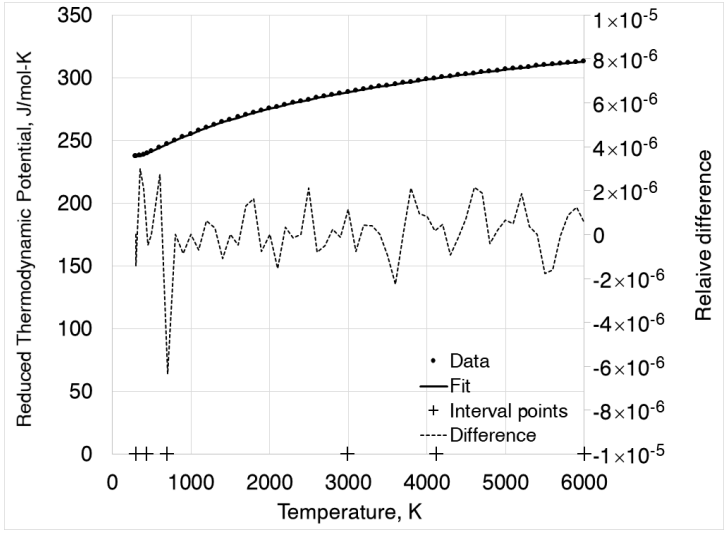
Beryllium Iodide (BeI) – Be1I1(g)

MW = 135.9167 g/mol – $\Delta_f H^o$(298.15 K) = 170.018 kJ/mol – S^o(298.15 K) = 237.297 J/mol-K – p^o = 0.1 MPa



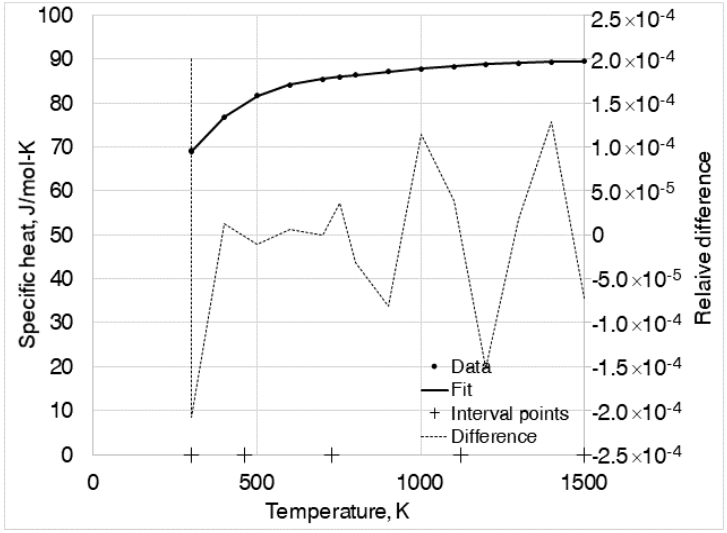


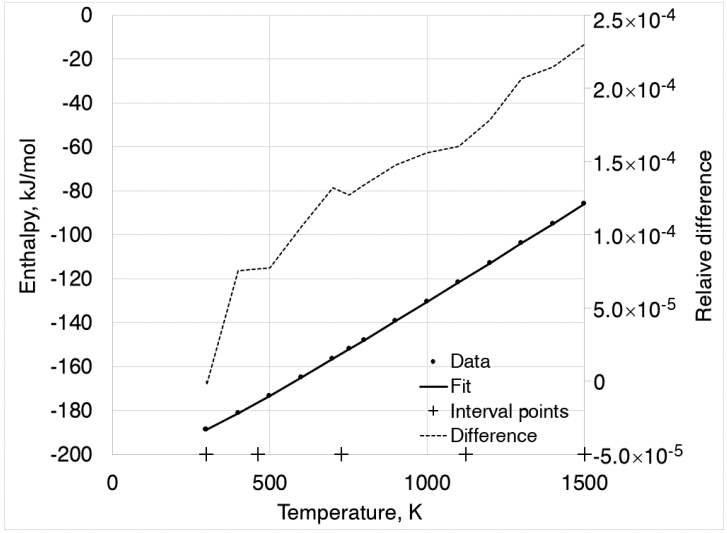


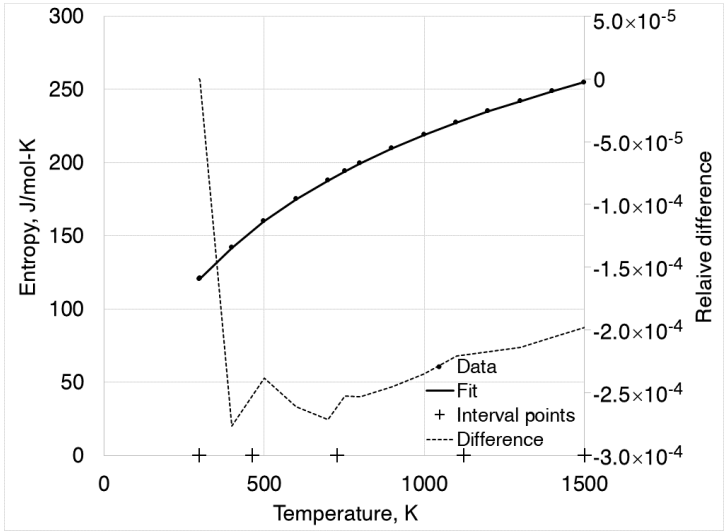


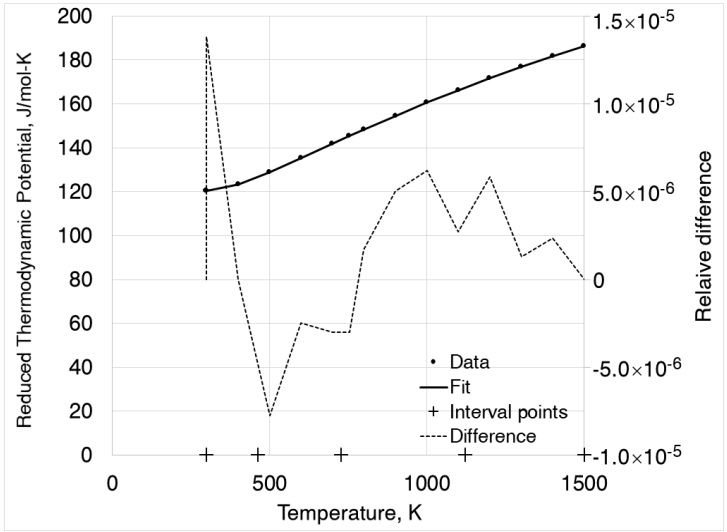
Beryllium Iodide (BeI_2) – Be1I2(cr)

MW = 262.8212 g/mol – $\Delta_f H^o$(298.15 K) = -188.698 kJ/mol – S^o(298.15 K) = 120.499 J/mol-K – p^o = 0.1 MPa



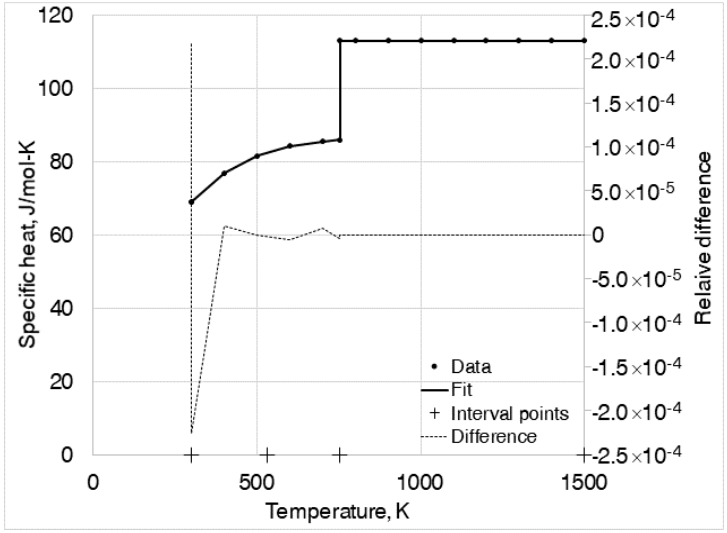


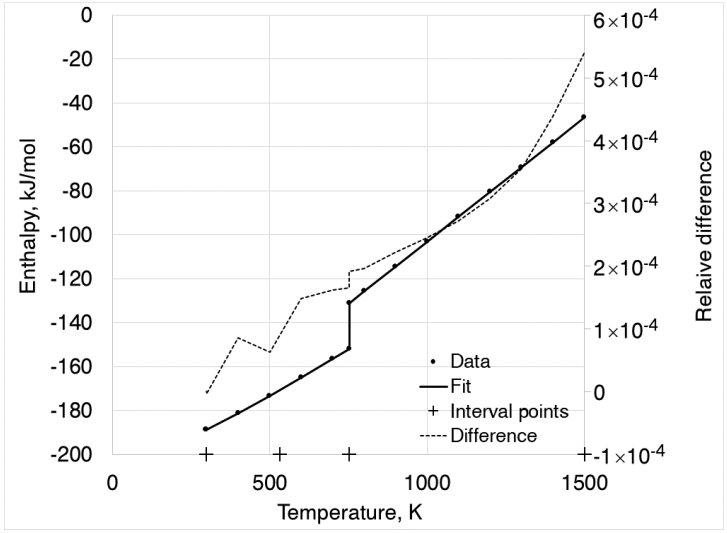


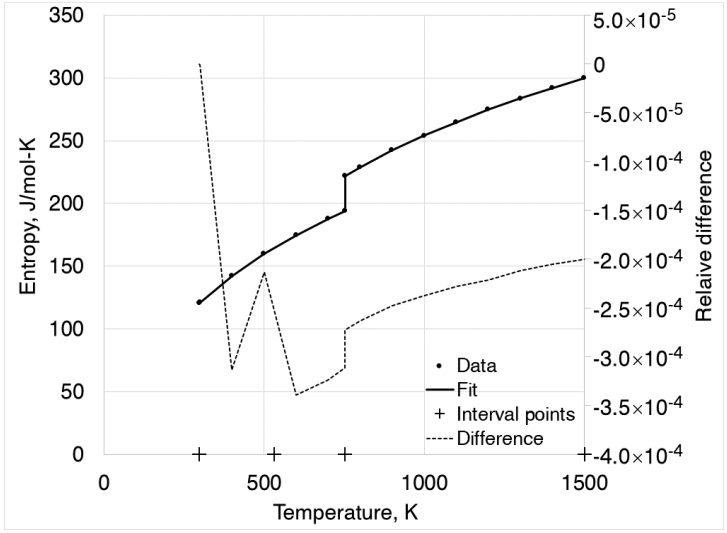


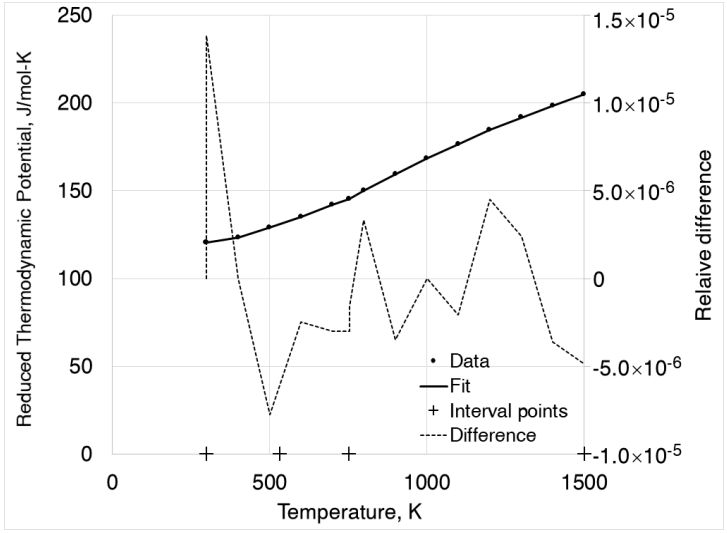
Beryllium Iodide (BeI_2) – Be1I2(cr,l)

MW = 262.8212 g/mol – $\Delta_f H^o$(298.15 K) = -188.698 kJ/mol – S^o(298.15 K) = 120.499 J/mol-K – p^o = 0.1 MPa



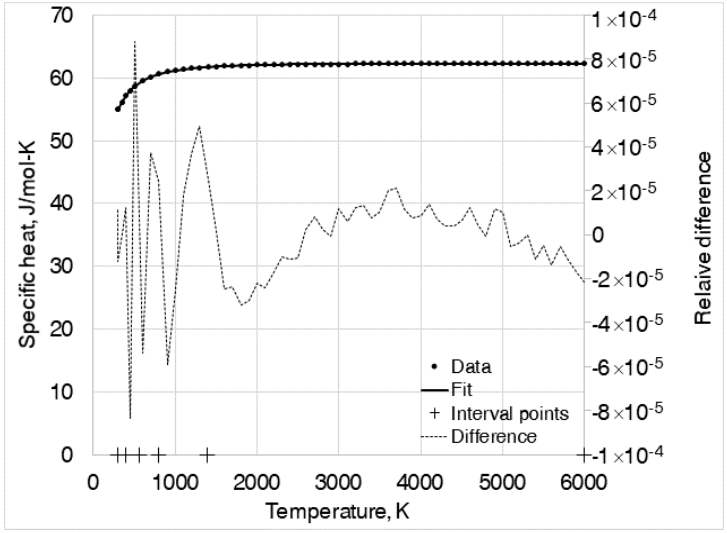


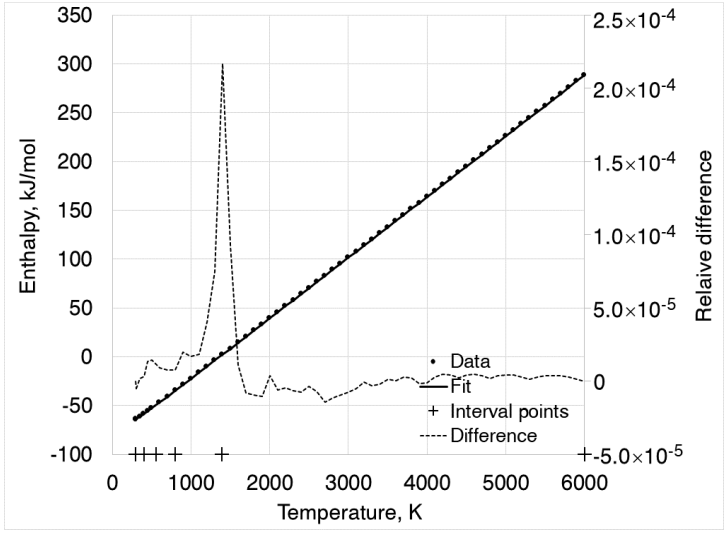


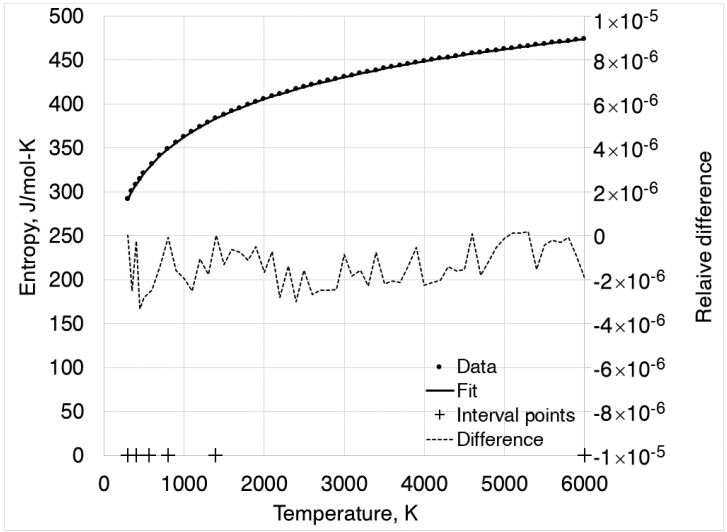


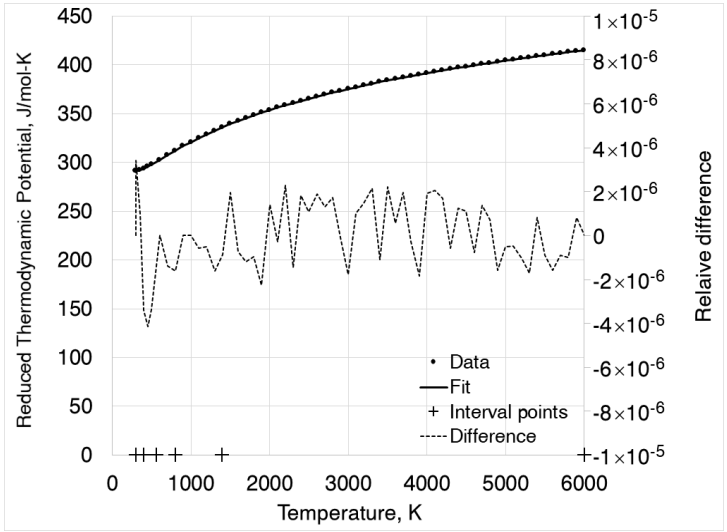
Beryllium Iodide (BeI_2) – Be1I2(g)

MW = 262.8212 g/mol – $\Delta_f H^o$(298.15 K) = -64.015 kJ/mol – S^o(298.15 K) = 291.524 J/mol-K – p^o = 0.1 MPa



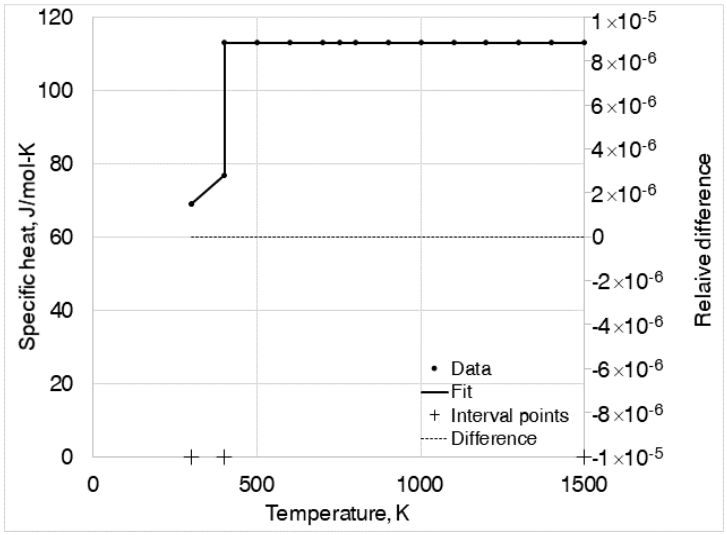


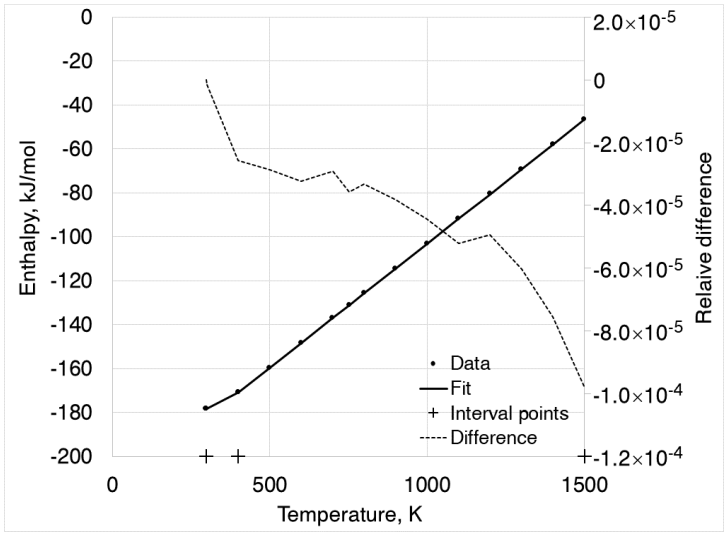


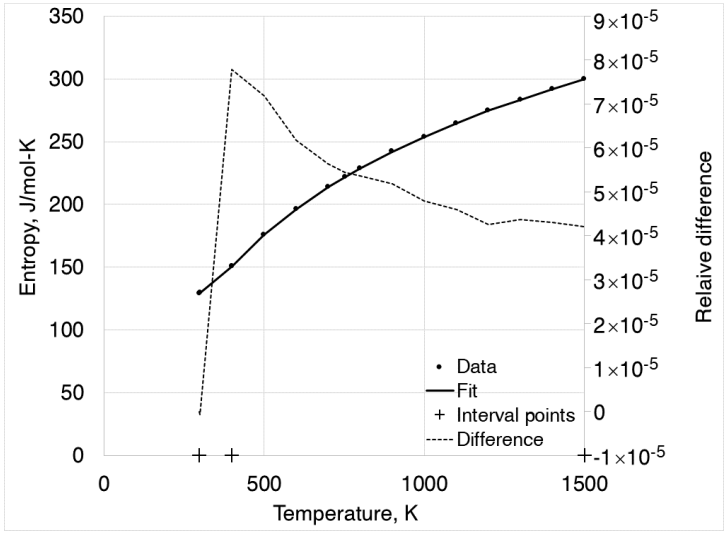


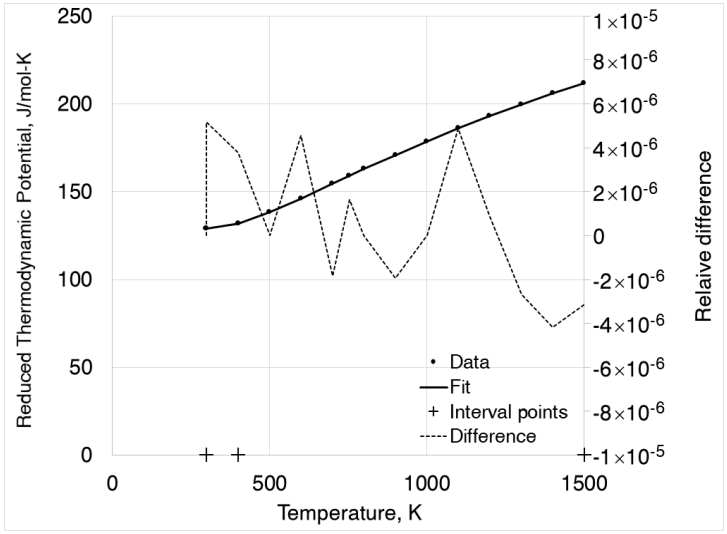
Beryllium Iodide (BeI_2) – Be1I2(l)

MW = 262.8212 g/mol – $\Delta_f H^o$(298.15 K) = -178.377 kJ/mol – S^o(298.15 K) = 128.999 J/mol-K – p^o = 0.1 MPa



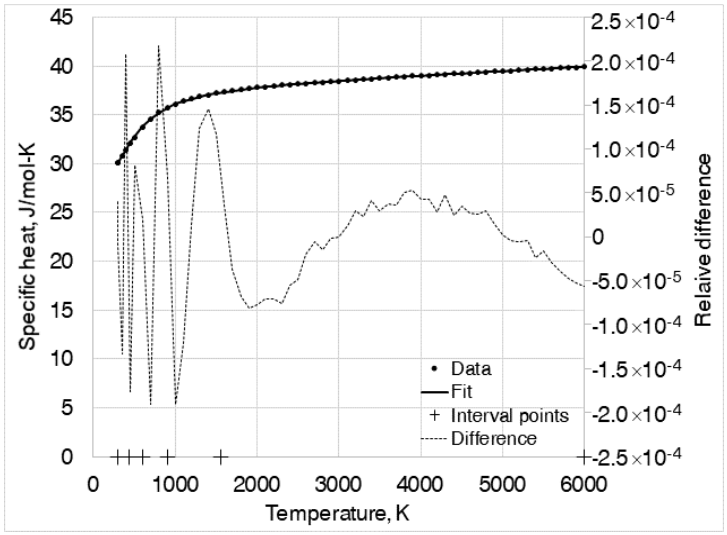


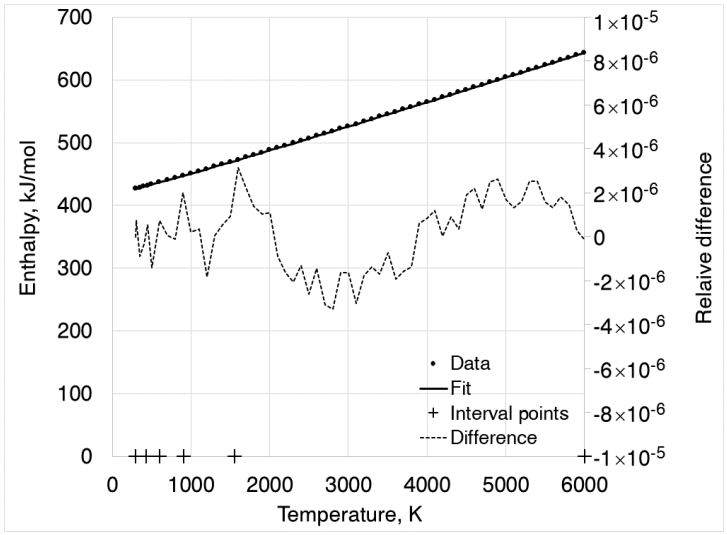


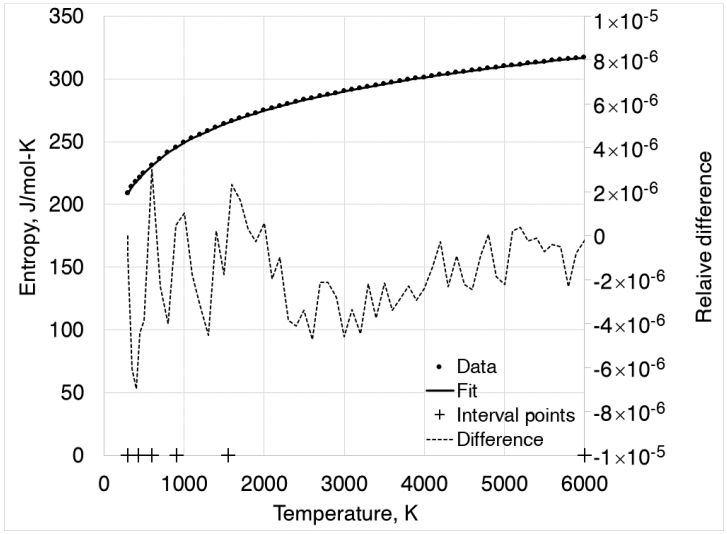


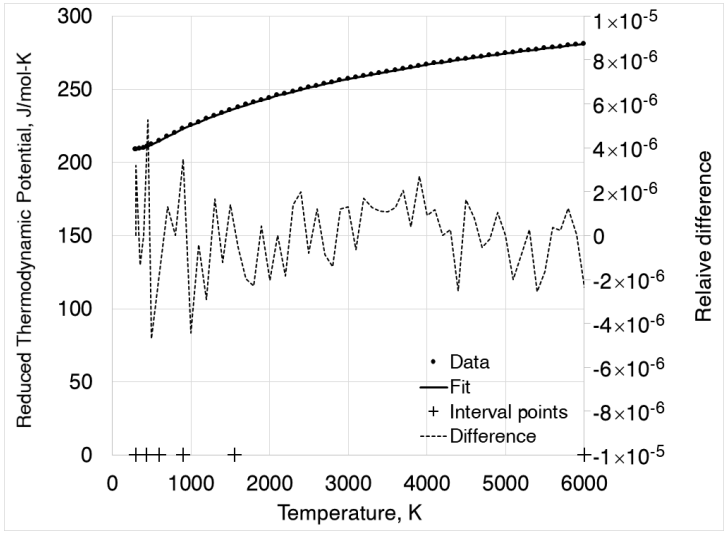
Beryllium Nitride (BeN) – Be1N1(g)

MW = 23.0189 g/mol – $\Delta_f H^o$(298.15 K) = 426.688 kJ/mol – S^o(298.15 K) = 208.768 J/mol-K – p^o = 0.1 MPa



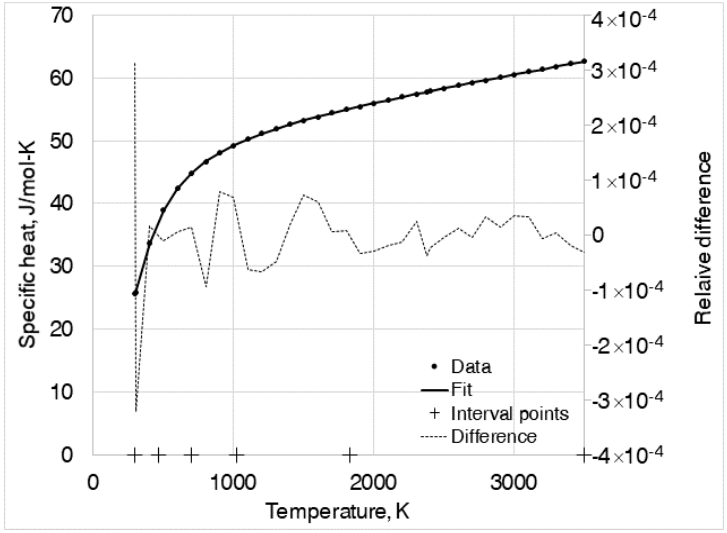


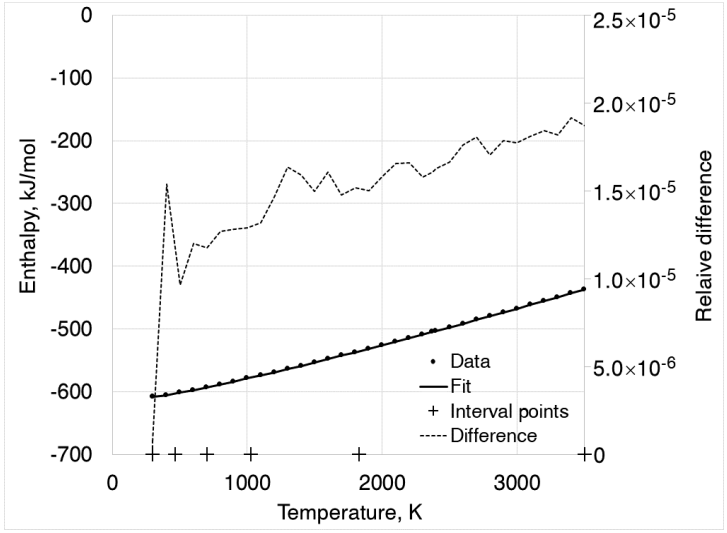


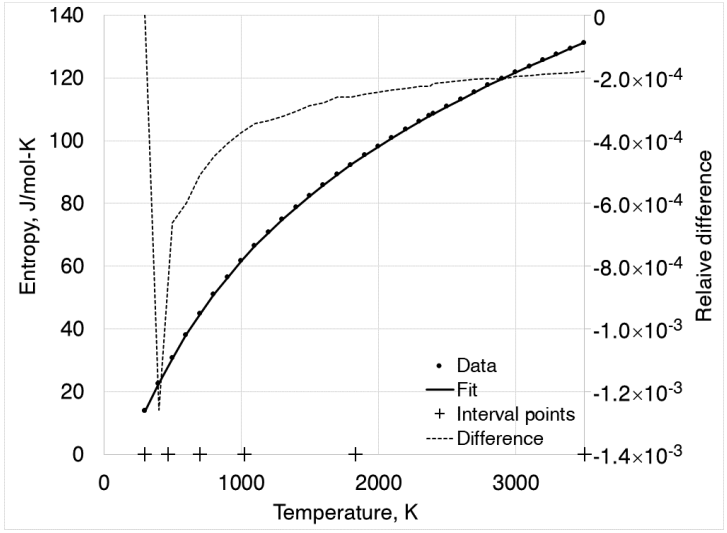


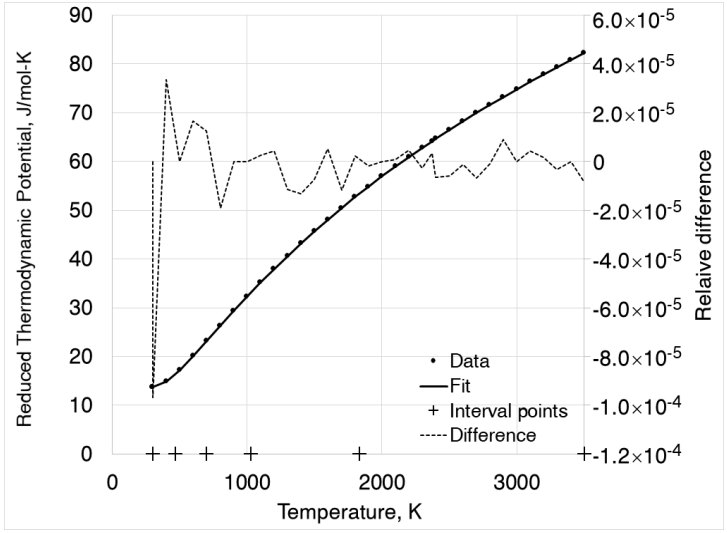
Beryllium Oxide, Alpha (BeO) – Be1O1(cr)_01

MW = 25.0116 g/mol – $\Delta_f H^o$(298.15 K) = -608.354 kJ/mol – S^o(298.15 K) = 13.77 J/mol-K – p^o = 0.1 MPa



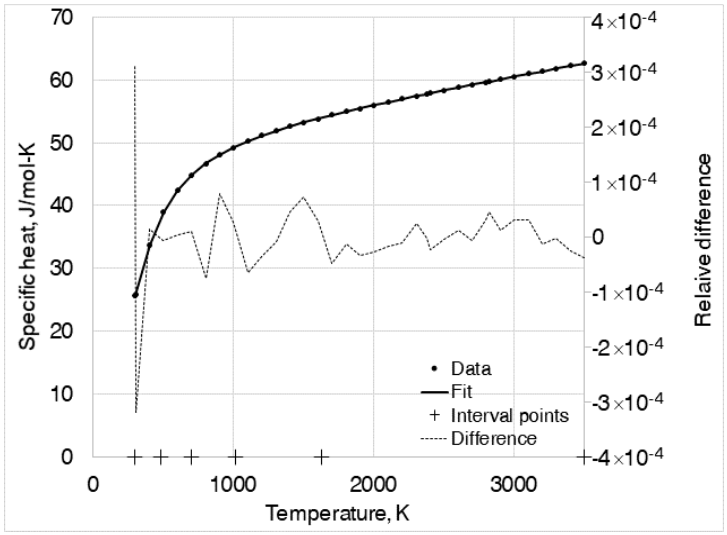


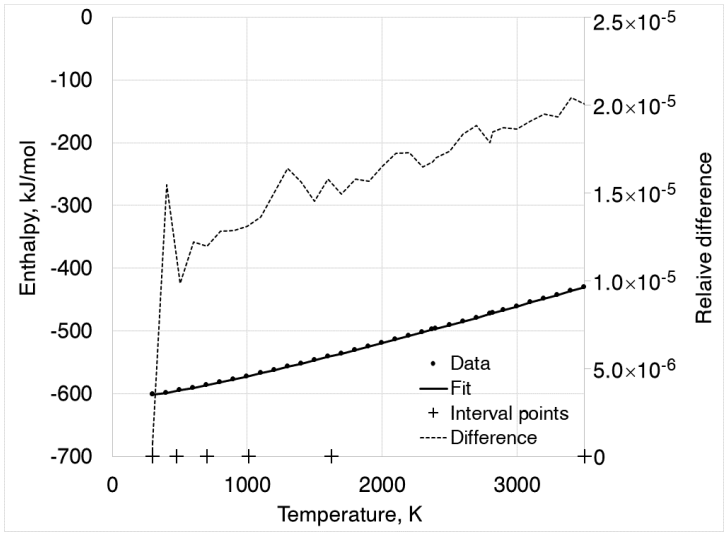


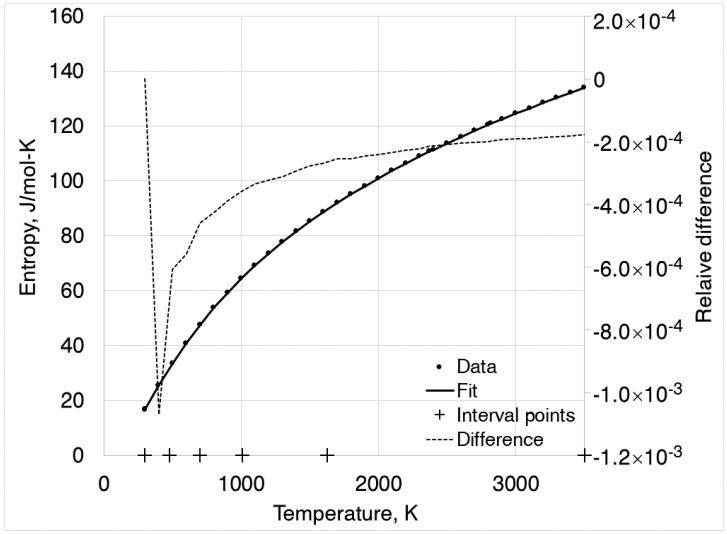


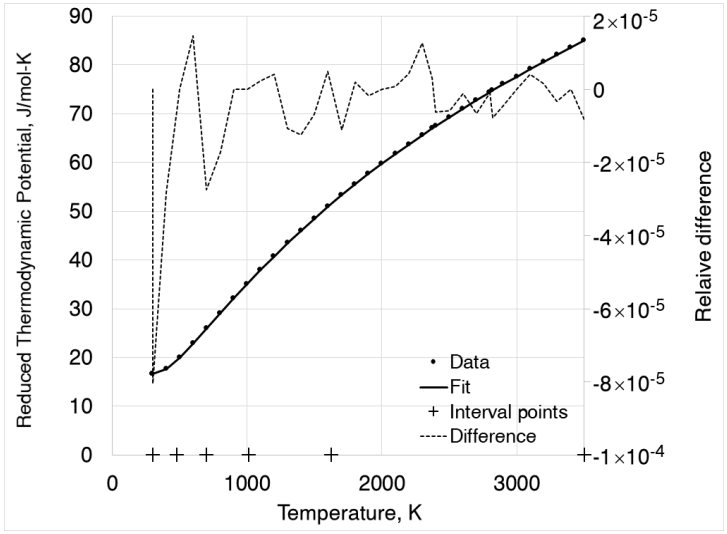
Beryllium Oxide, Beta (BeO) – Be1O1(cr)_02

MW = 25.0116 g/mol – $\Delta_f H^o$(298.15 K) = -601.659 kJ/mol – S^o(298.15 K) = 16.591 J/mol-K – p^o = 0.1 MPa



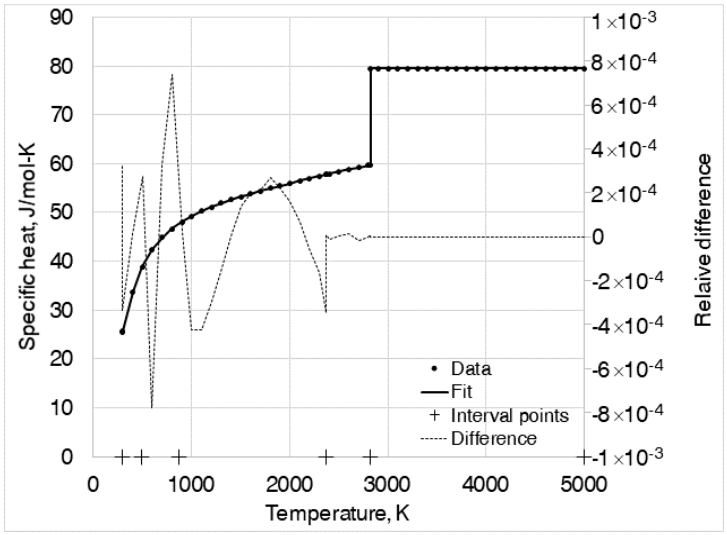


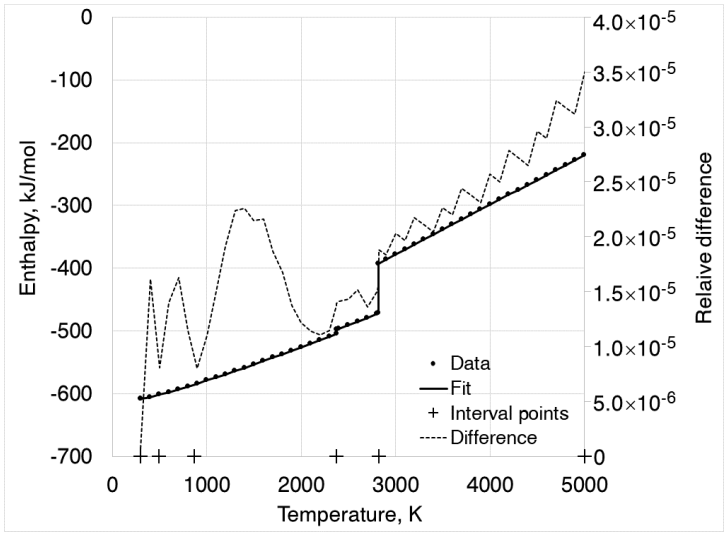


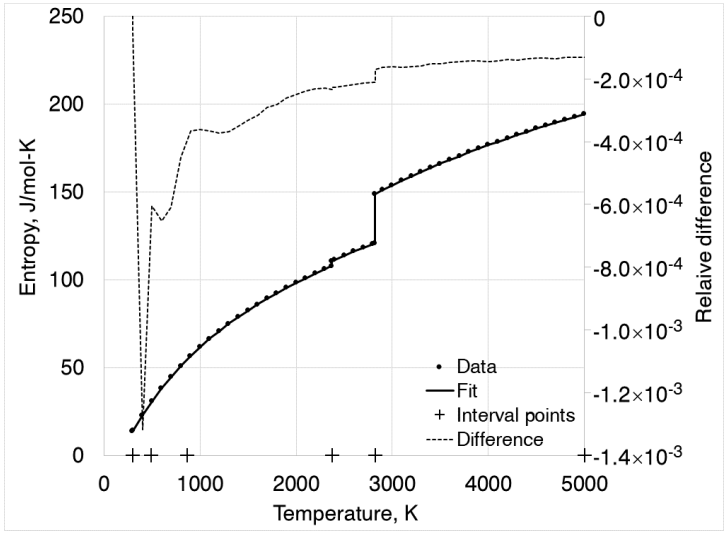


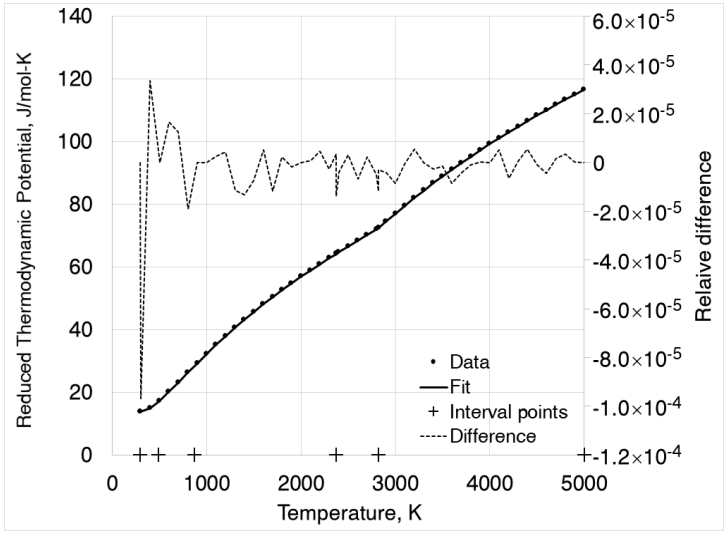
Beryllium Oxide (BeO) – Be1O1(cr,l)

MW = 25.0116 g/mol – $\Delta_f H^o$(298.15 K) = -608.354 kJ/mol – S^o(298.15 K) = 13.77 J/mol-K – p^o = 0.1 MPa



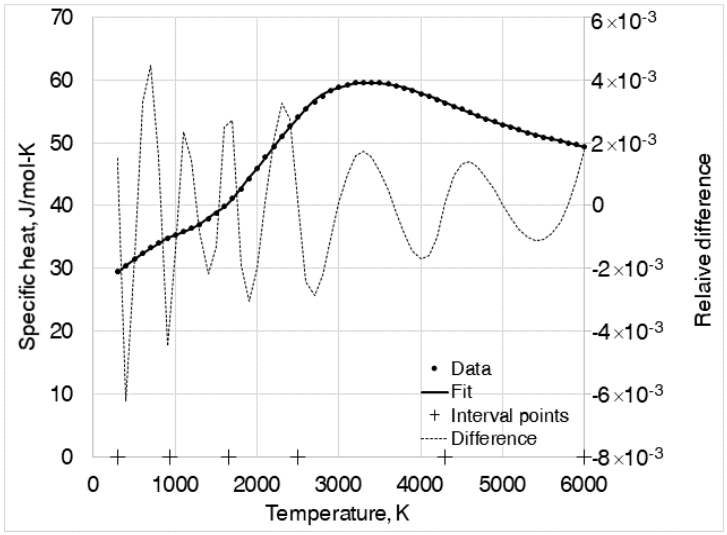


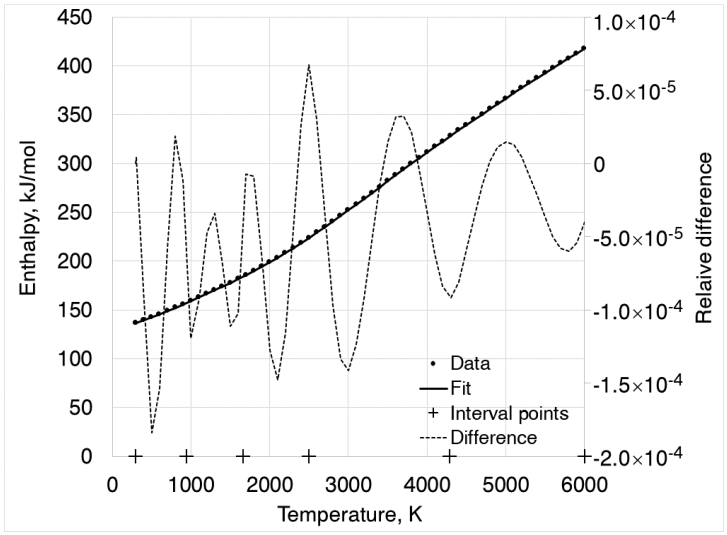


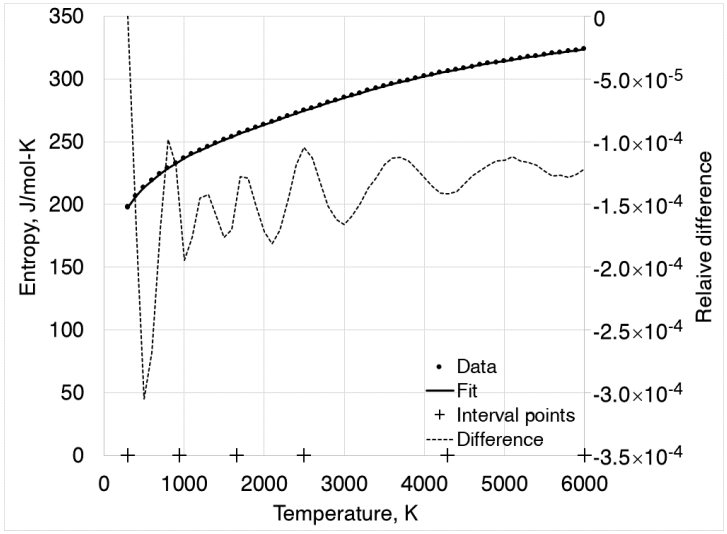


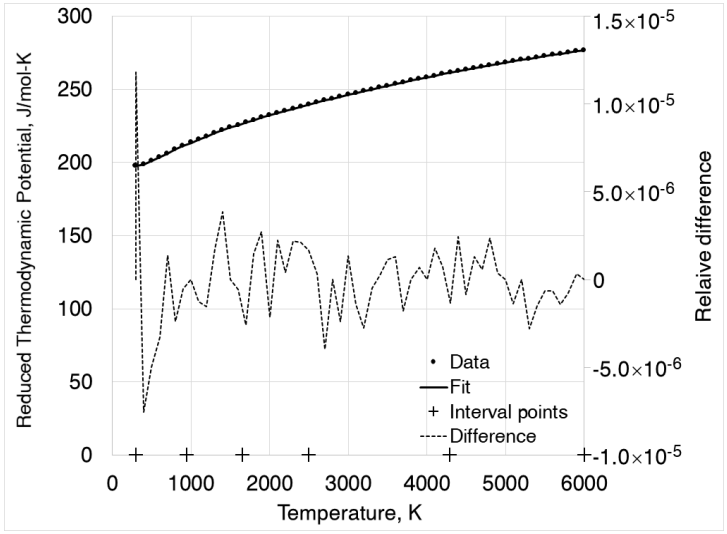
Beryllium Oxide (BeO) – Be1O1(g)

MW = 25.0116 g/mol – $\Delta_f H^o$(298.15 K) = 136.398 kJ/mol – S^o(298.15 K) = 197.625 J/mol-K – p^o = 0.1 MPa



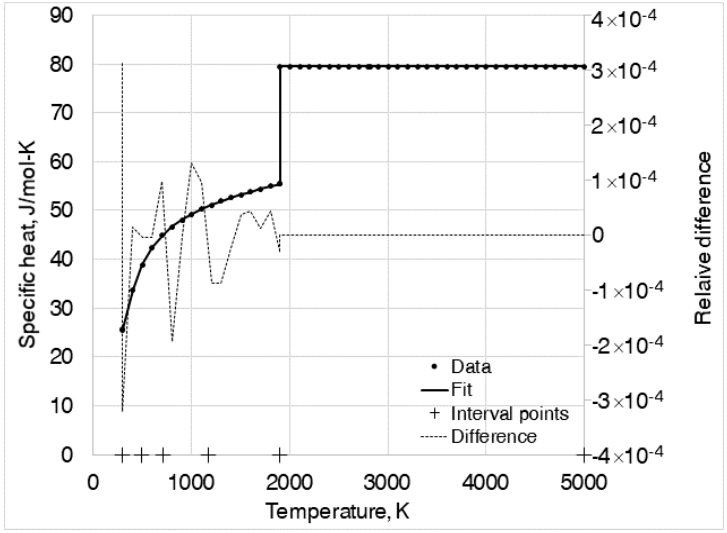


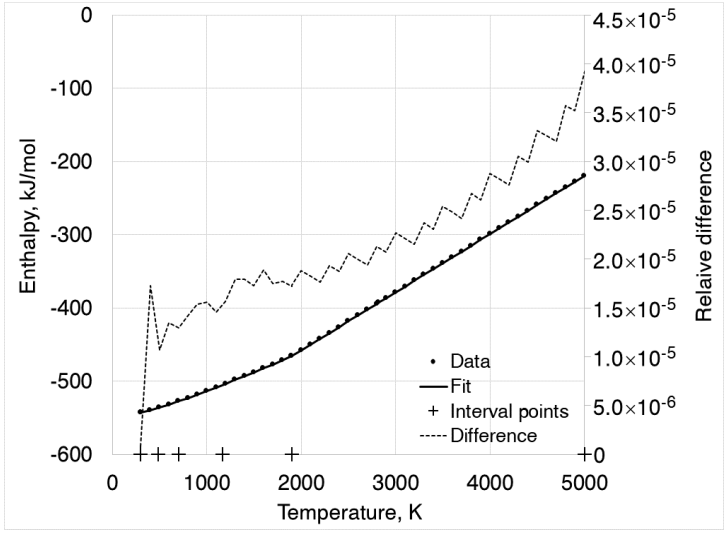


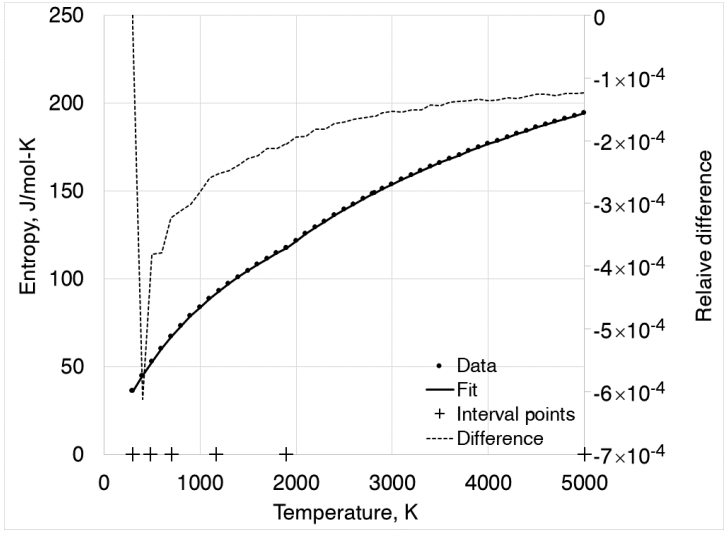


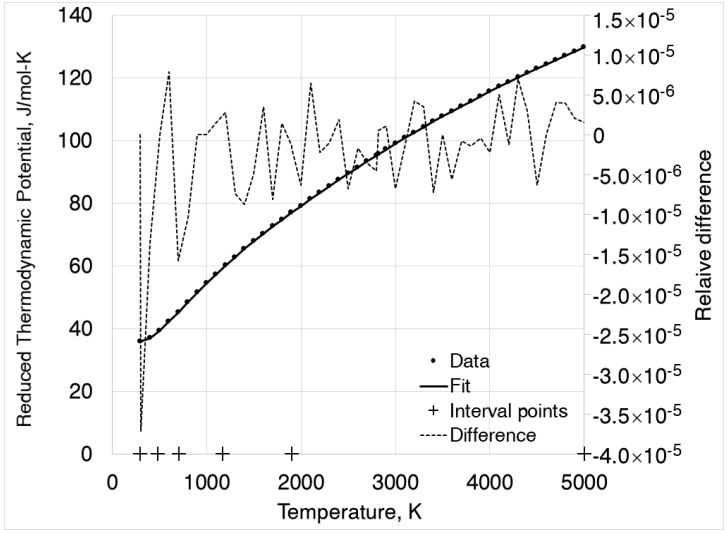
Beryllium Oxide (BeO) – Be1O1(l)

MW = 25.0116 g/mol – $\Delta_f H^o$(298.15 K) = -542.705 kJ/mol – S^o(298.15 K) = 35.931 J/mol-K – p^o = 0.1 MPa



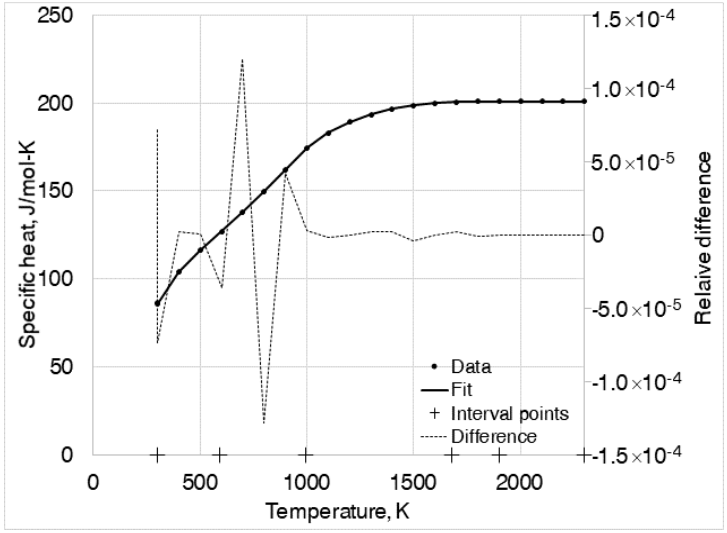


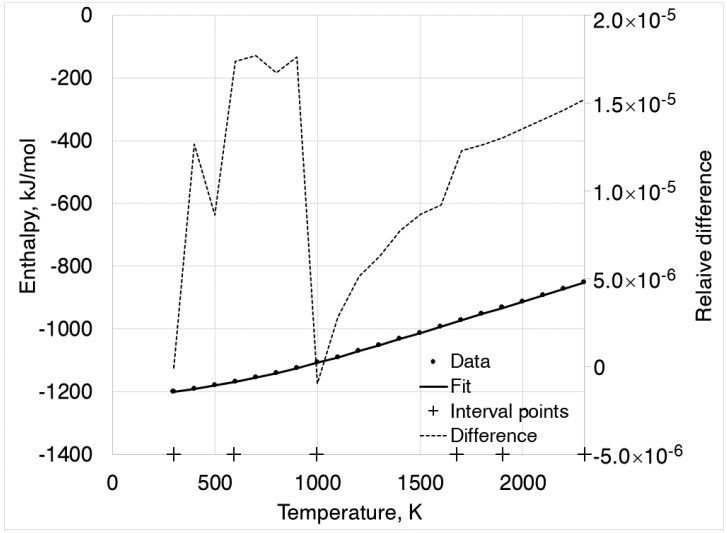


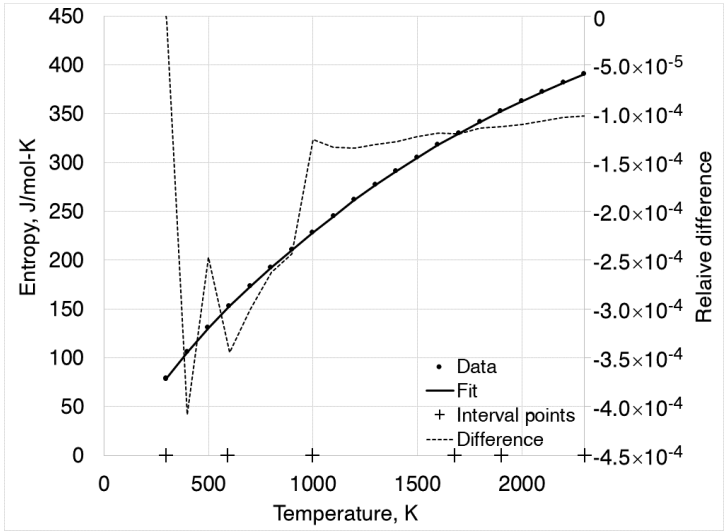


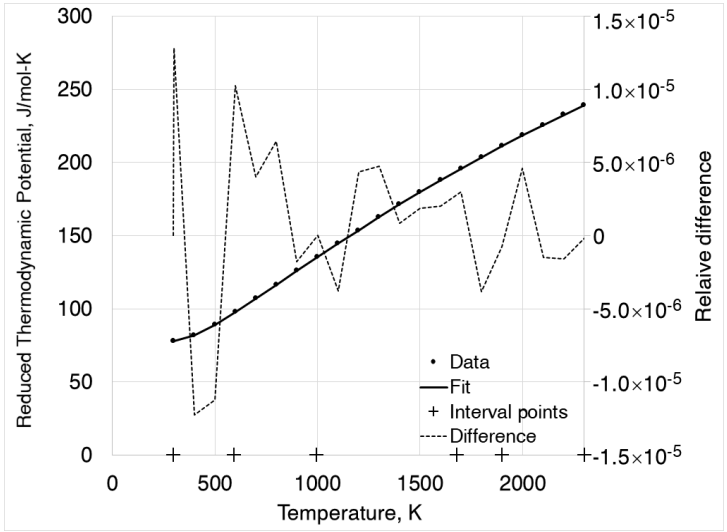
Beryllium Sulfate, Alpha ($BeSO_4$) – Be1O4S1(cr)_01

MW = 105.07 g/mol – $\Delta_f H^o$(298.15 K) = -1200.808 kJ/mol – S^o(298.15 K) = 77.971 J/mol-K – p^o = 0.1 MPa



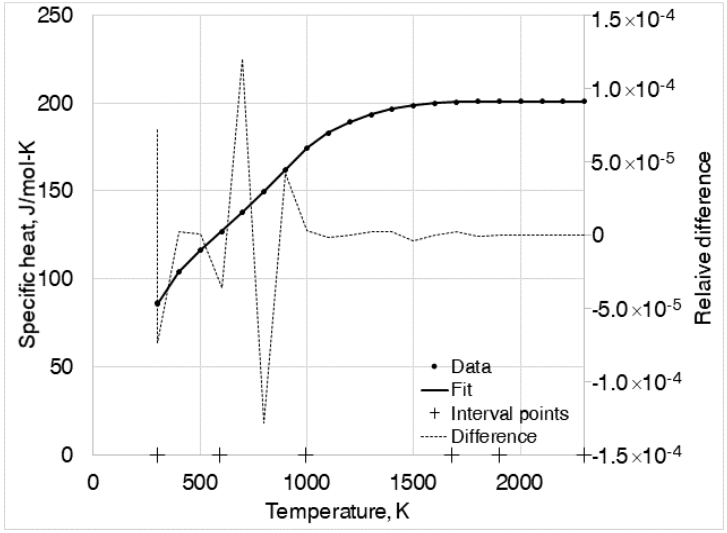


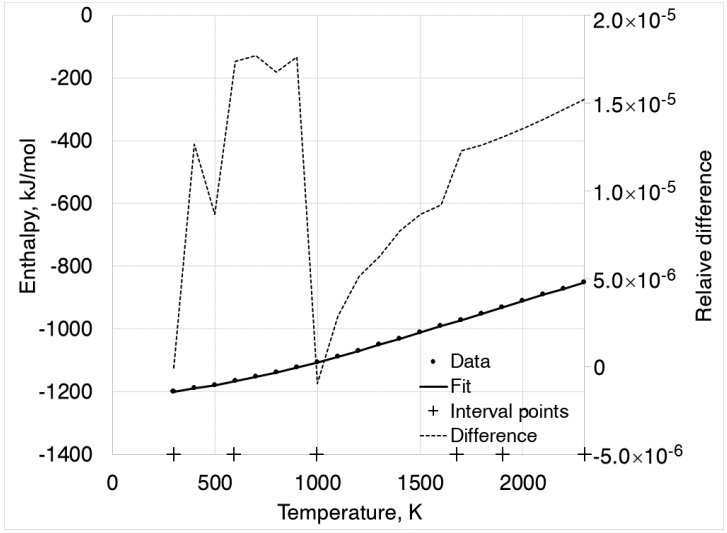


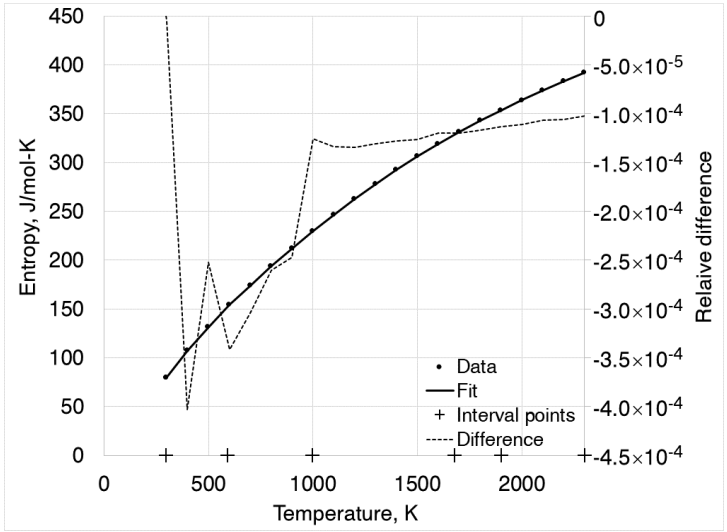


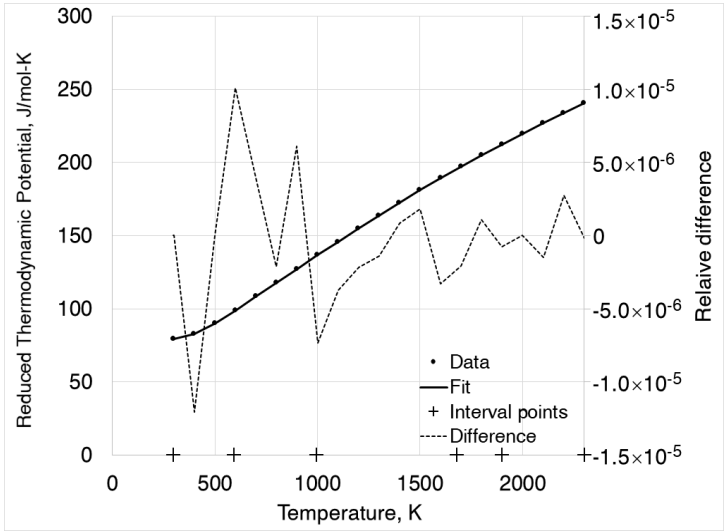
Beryllium Sulfate, Beta ($BeSO_4$) – Be1O4S1(cr)_02

MW = 105.07 g/mol – $\Delta_f H^o$(298.15 K) = -1199.695 kJ/mol – S^o(298.15 K) = 79.258 J/mol-K – p^o = 0.1 MPa



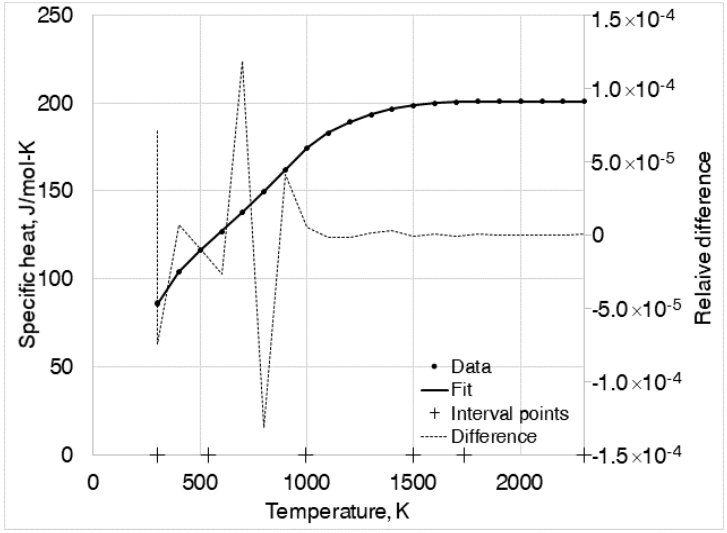


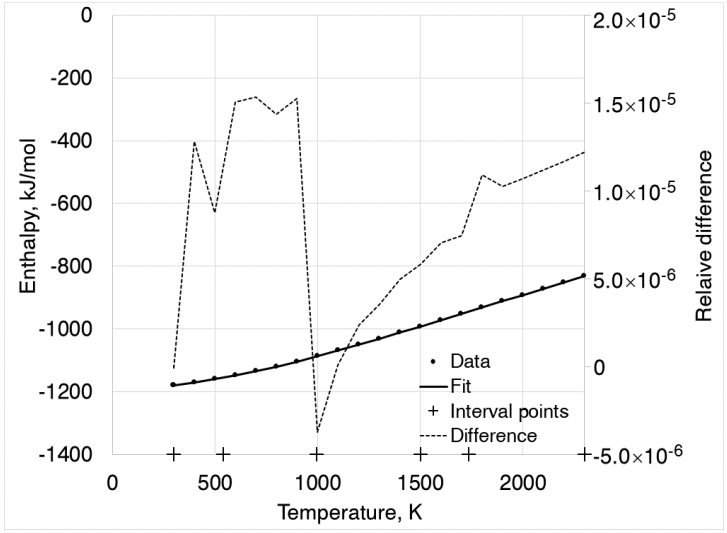


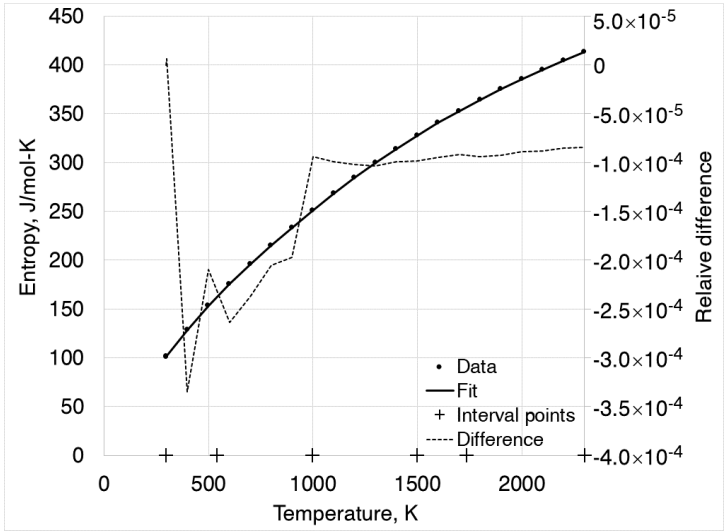


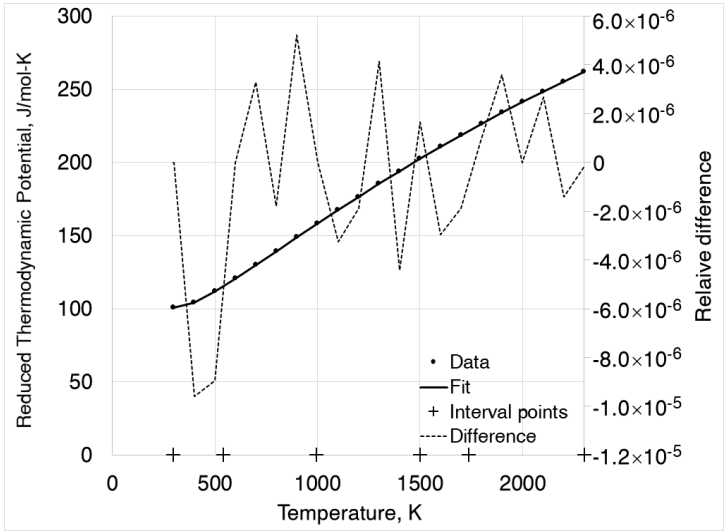
Beryllium Sulfate, Gamma ($BeSO_4$) – Be1O4S1(cr)_03

MW = 105.07 g/mol – $\Delta_f H^o$(298.15 K) = -1180.143 kJ/mol – S^o(298.15 K) = 100.784 J/mol-K – p^o = 0.1 MPa



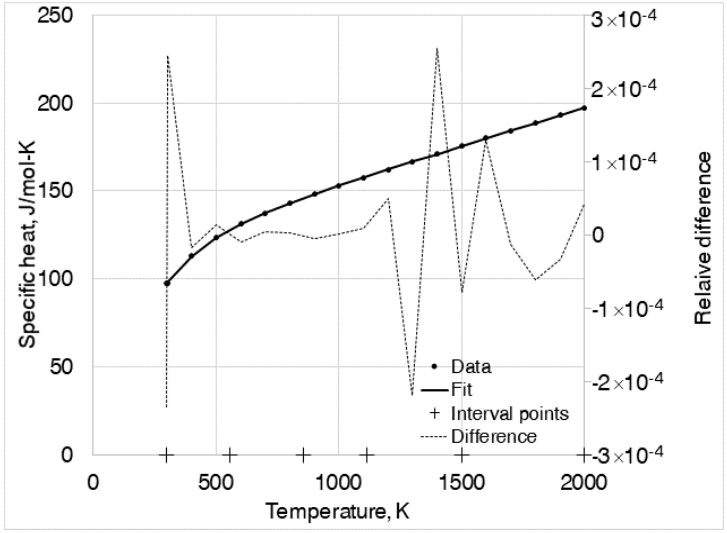


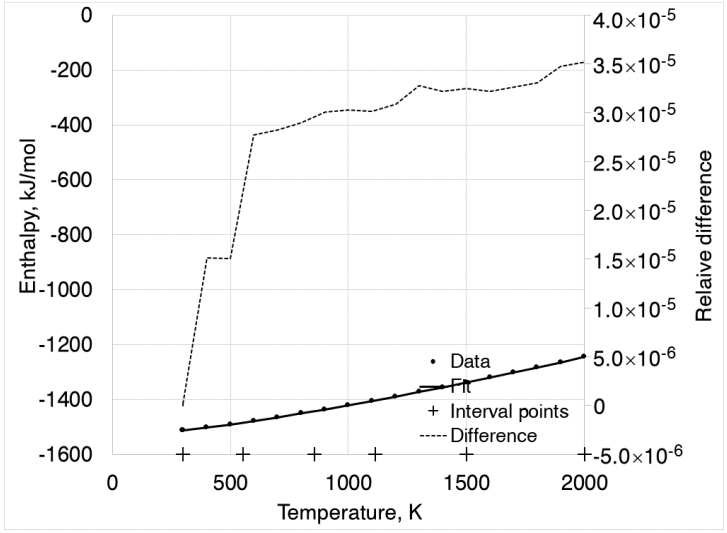


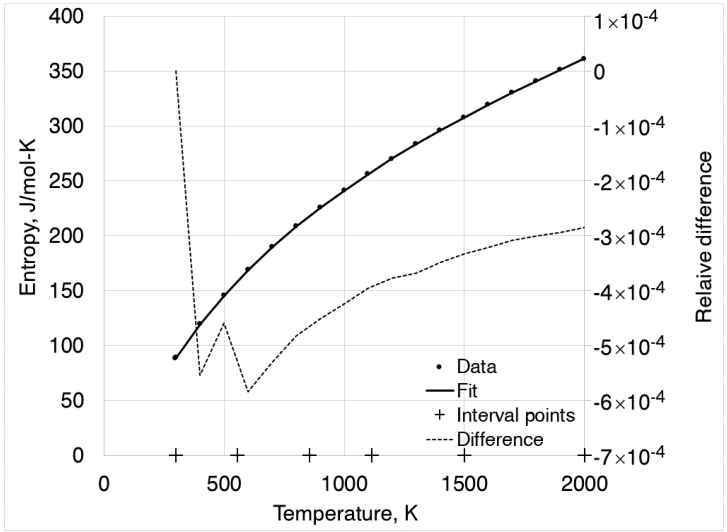


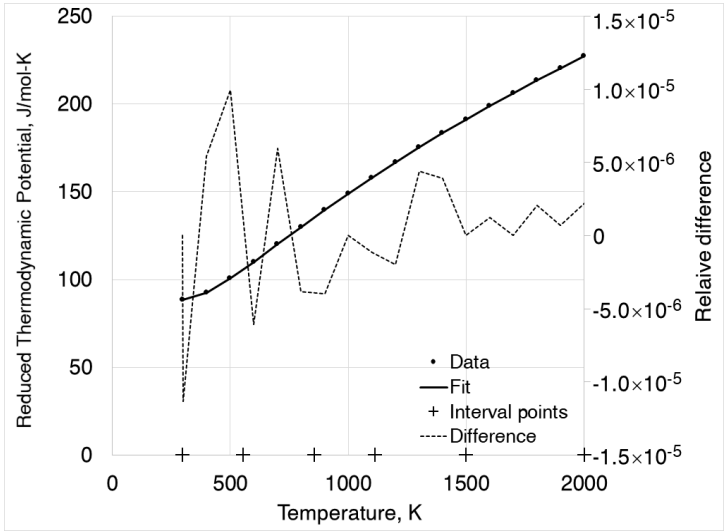
Beryllium Tungsten Oxide ($BeWO_4$) – Be1O4W1(cr)

MW = 256.86 g/mol – $\Delta_f H^o$(298.15 K) = -1513.353 kJ/mol – S^o(298.15 K) = 88.366 J/mol-K – p^o = 0.1 MPa



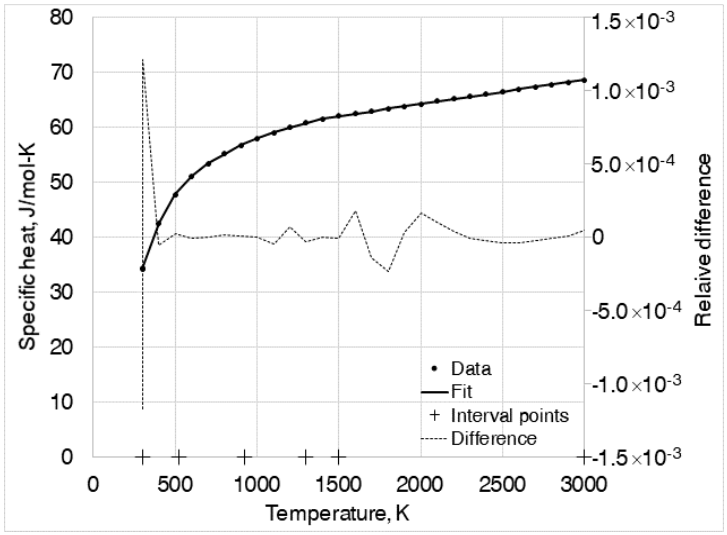


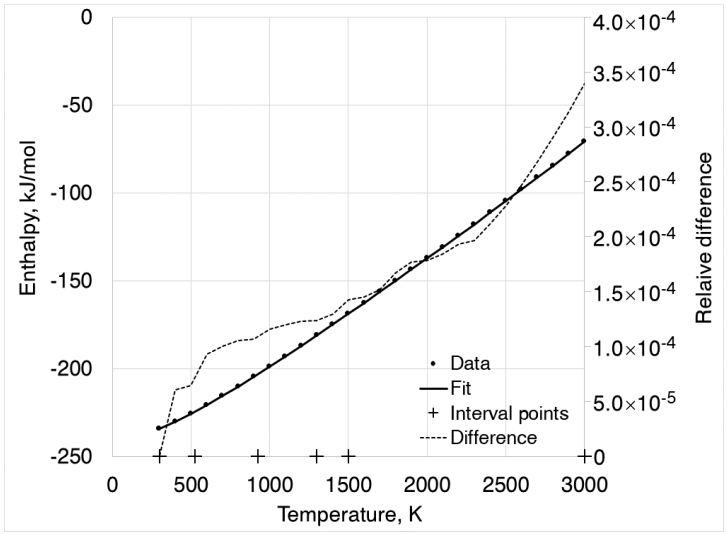


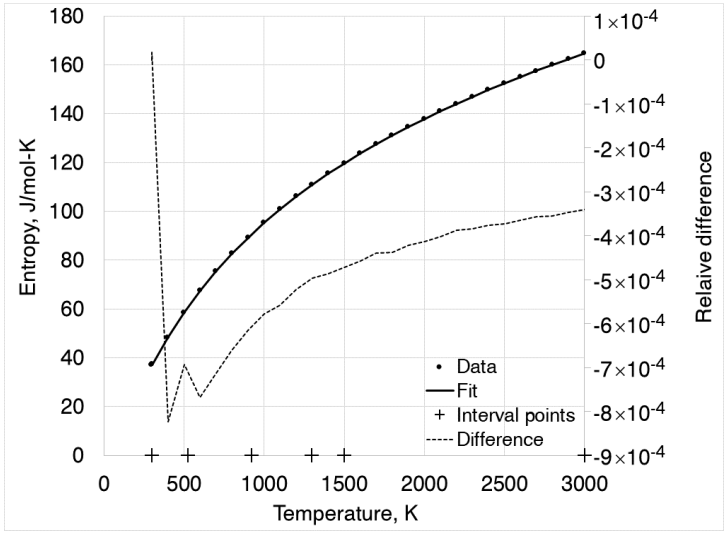


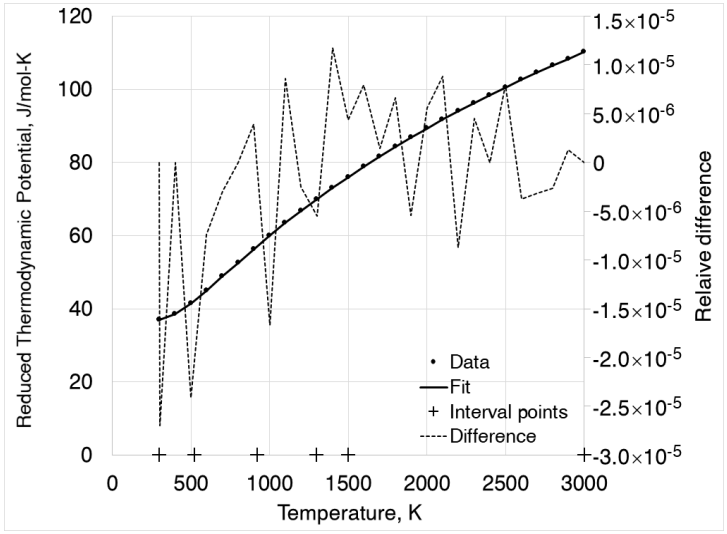
Beryllium Sulfide (BeS) – Be1S1(cr)

MW = 41.07 g/mol – $\Delta_f H^o$(298.15 K) = -234.304 kJ/mol – S^o(298.15 K) = 37.028 J/mol-K – p^o = 0.1 MPa



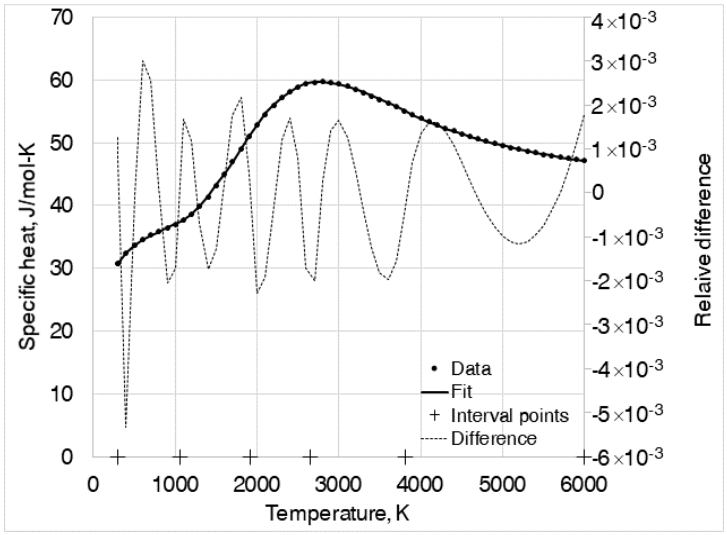


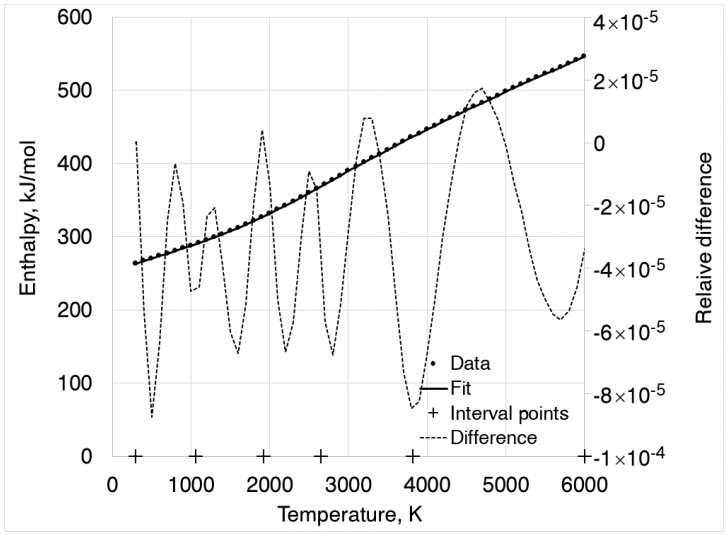


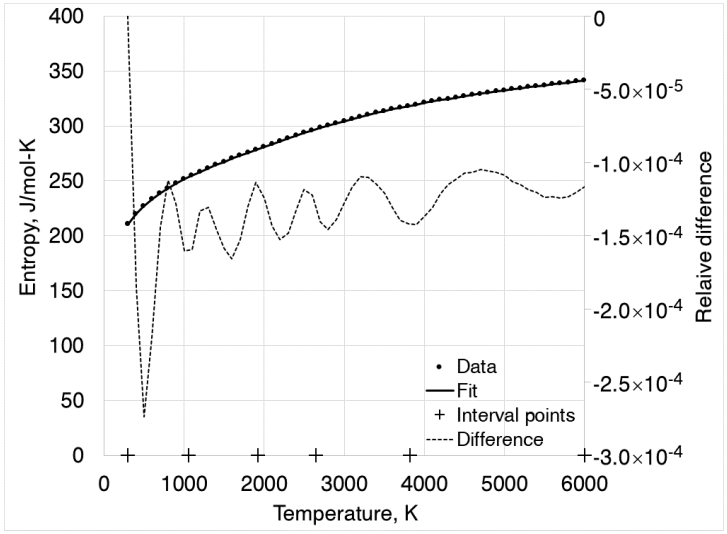


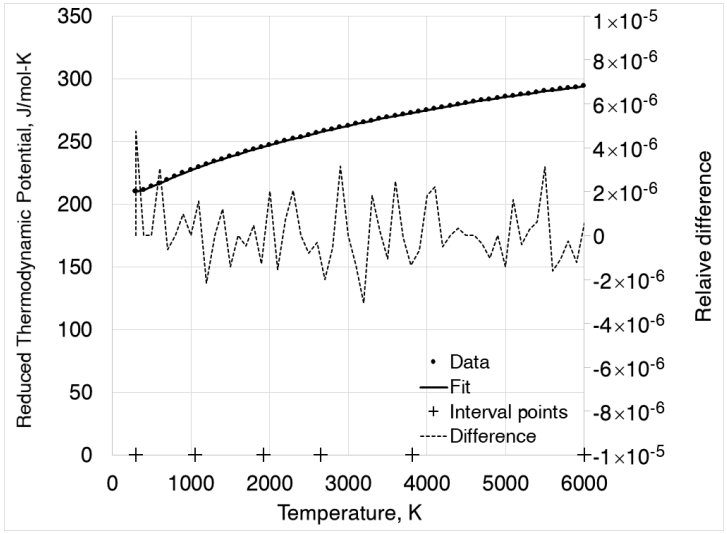
Beryllium Sulfide (BeS) – Be1S1(g)

MW = 41.07 g/mol – $\Delta_f H^o$(298.15 K) = 263.592 kJ/mol – S^o(298.15 K) = 210.288 J/mol-K – p^o = 0.1 MPa



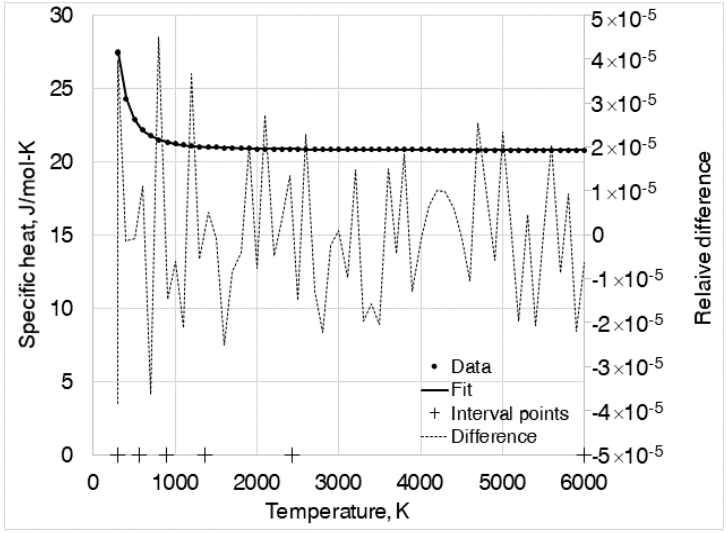


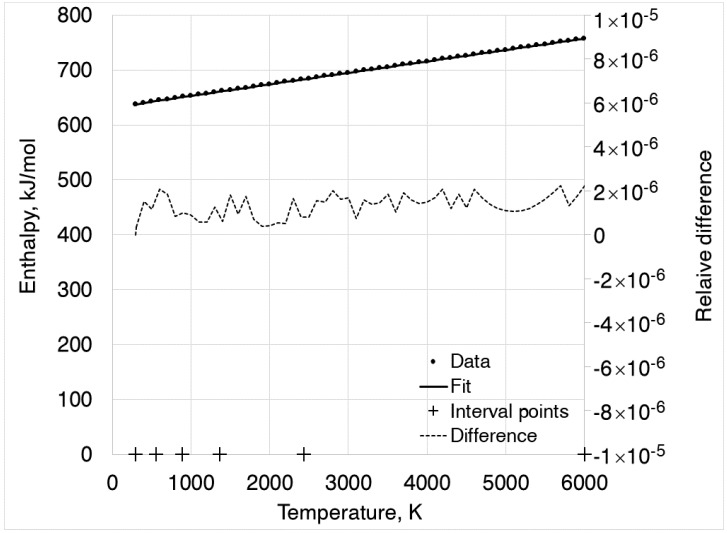


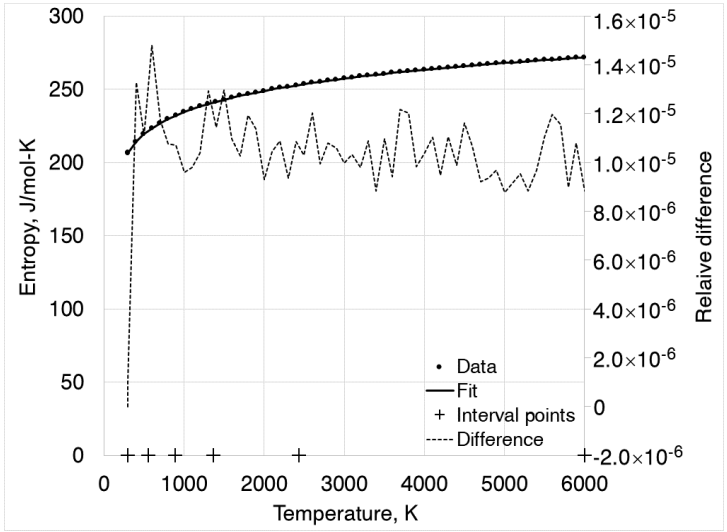


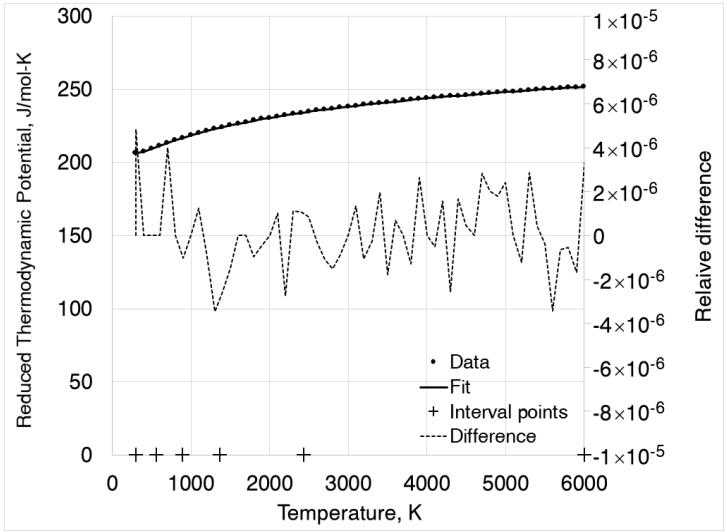
Beryllium (Be_2) – Be2(g)

MW = 18.02436 g/mol – $\Delta_f H^o$(298.15 K) = 637.161 kJ/mol – S^o(298.15 K) = 206.41 J/mol-K – p^o = 0.1 MPa



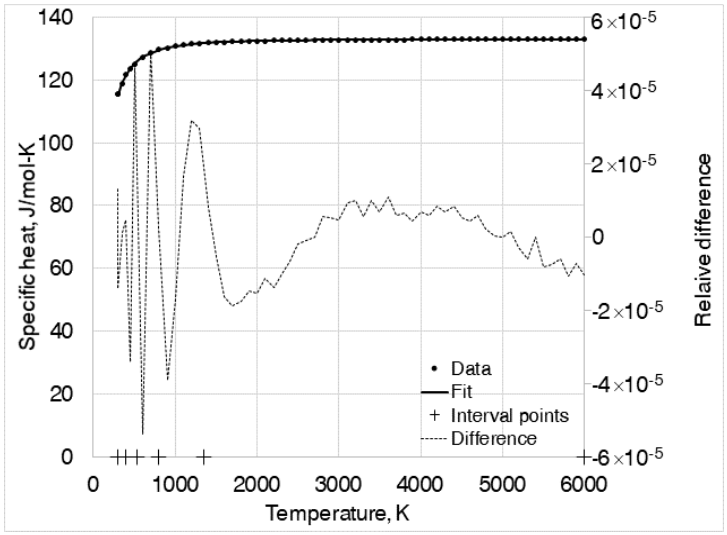


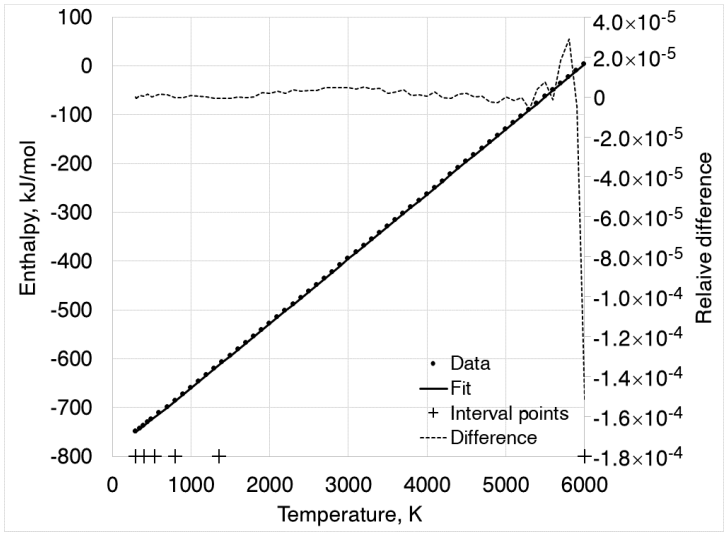


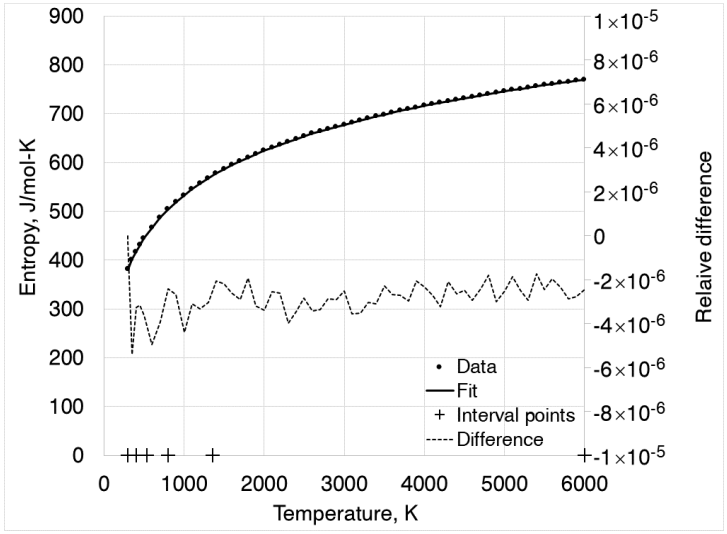


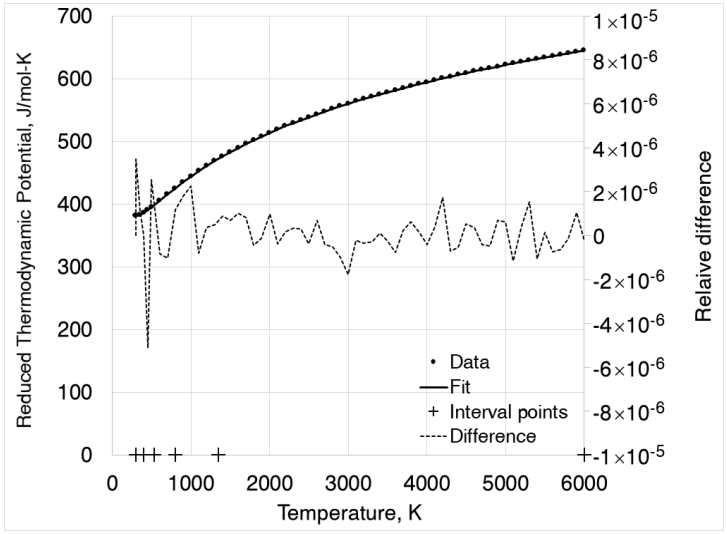
Beryllium Chloride ($(BeCl_2)_2$) – Be2Cl4(g)

MW = 159.836 g/mol – $\Delta_f H^o$(298.15 K) = -748.936 kJ/mol – S^o(298.15 K) = 381.475 J/mol-K – p^o = 0.1 MPa



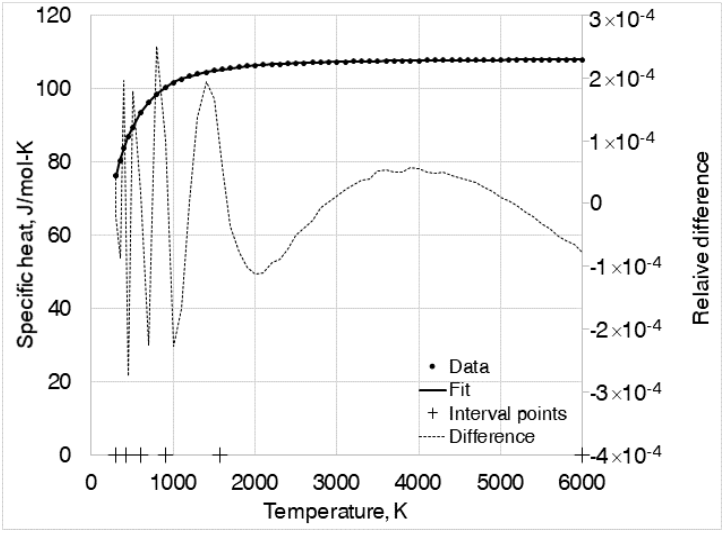


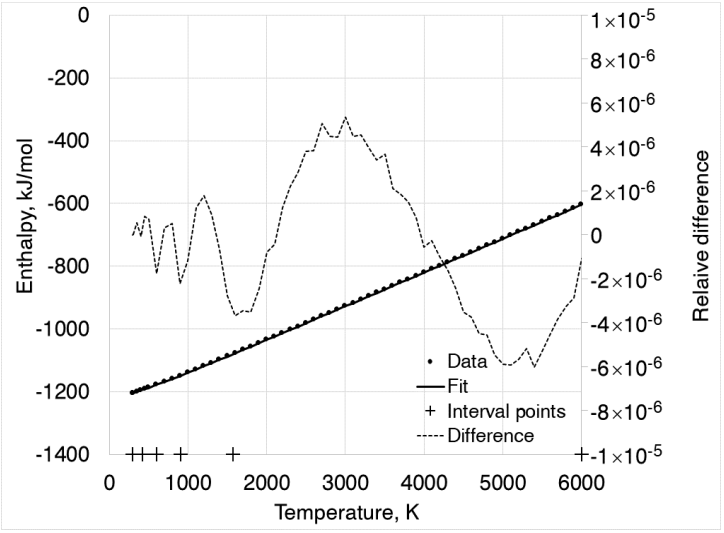


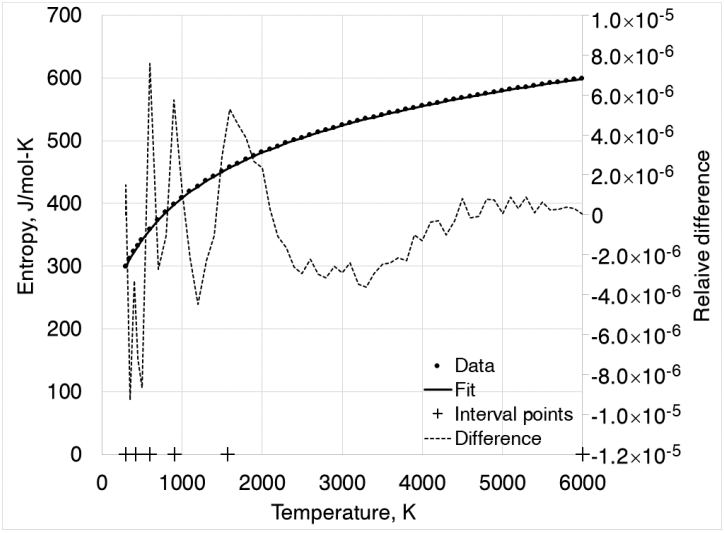


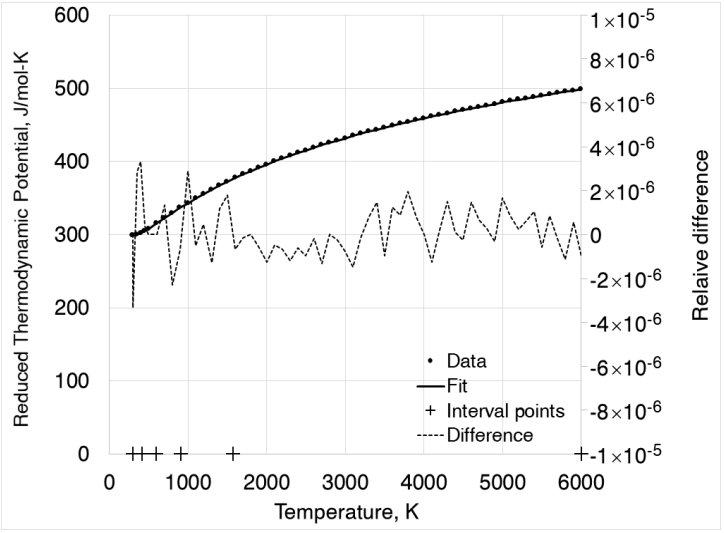
Beryllium Fluoride Oxide ($O(BeF)_2$) – Be2F2O1(g)

MW = 72.0206 g/mol – $\Delta_f H^o$(298.15 K) = -1204.574 kJ/mol – S^o(298.15 K) = 298.915 J/mol-K – p^o = 0.1 MPa



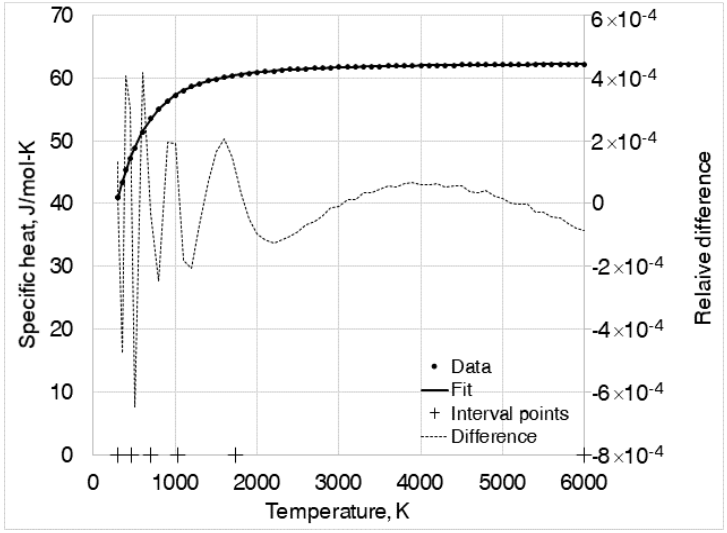


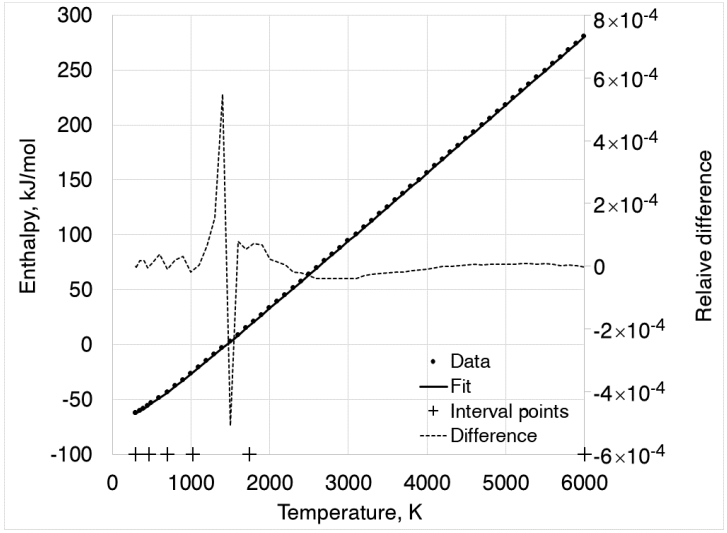


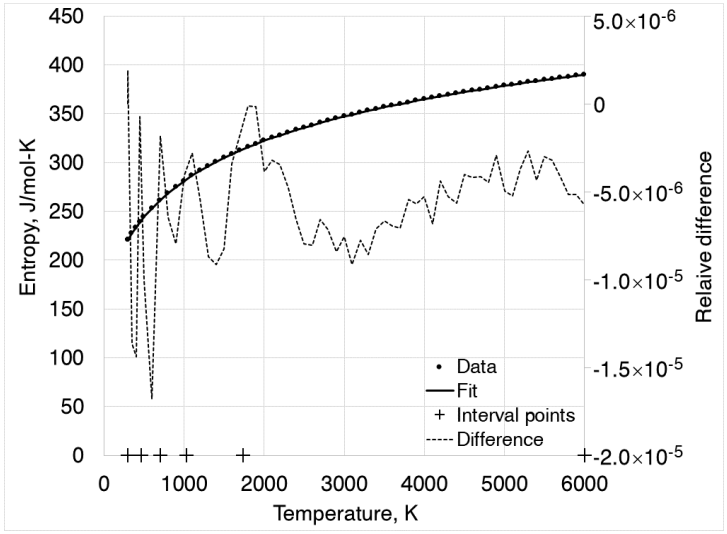


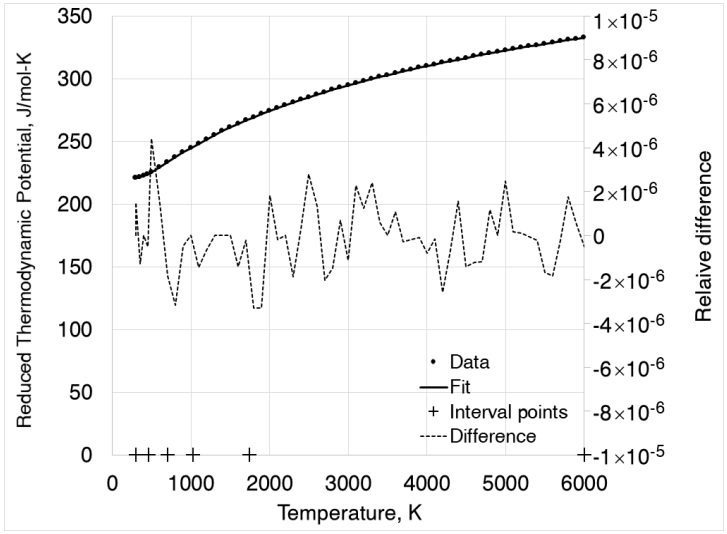
Beryllium Oxide (Be_2O) – Be2O1(g)

MW = 34.0238 g/mol – $\Delta_f H^o$(298.15 K) = -62.76 kJ/mol – S^o(298.15 K) = 220.827 J/mol-K – p^o = 0.1 MPa



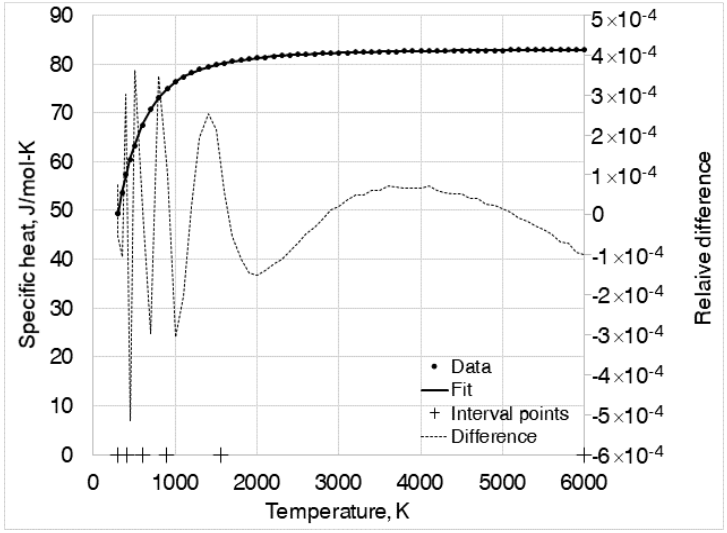


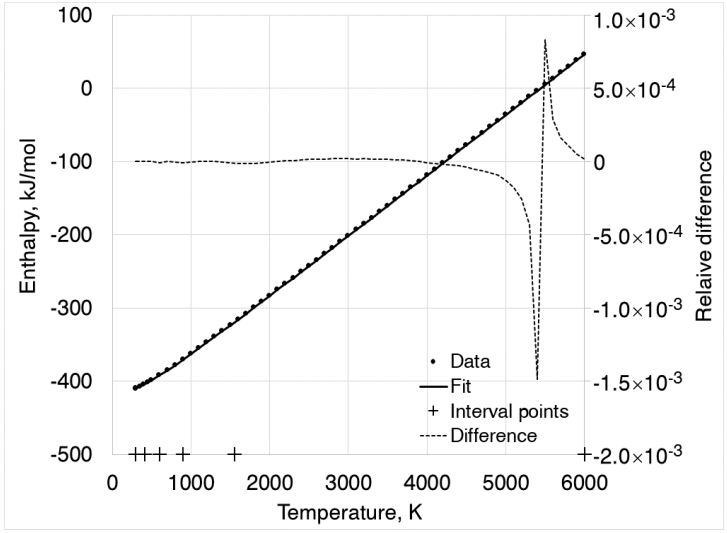


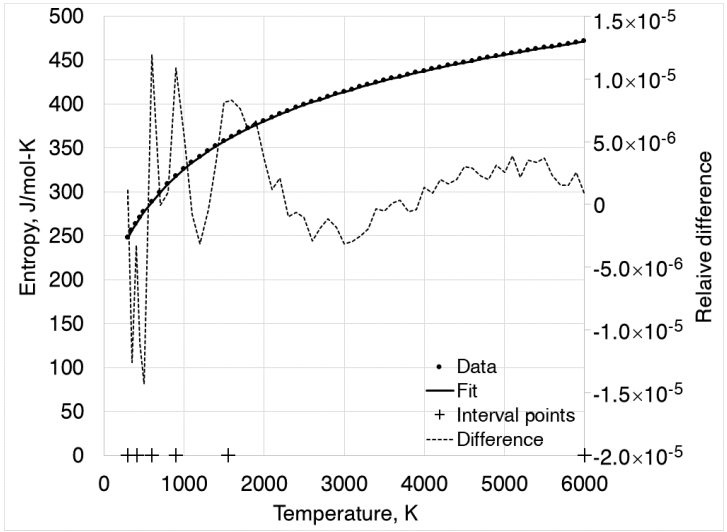


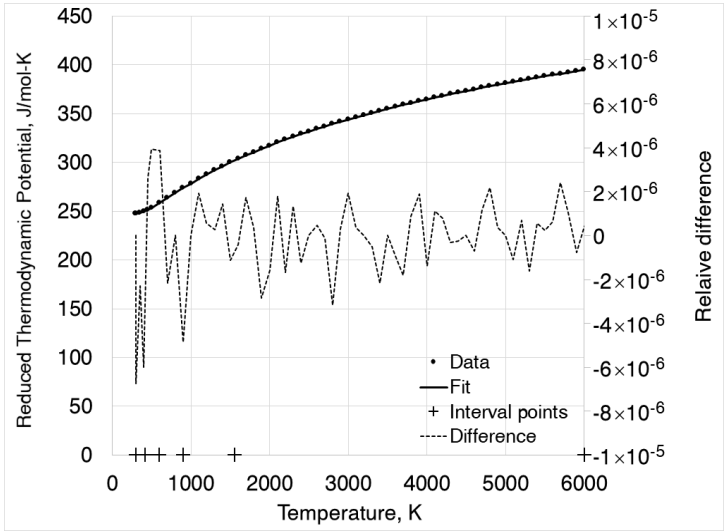
Beryllium Oxide ((BeO)$_2$) – Be2O2(g)

MW = 50.0232 g/mol – $\Delta_f H^o$(298.15 K) = -410.032 kJ/mol – S^o(298.15 K) = 247.611 J/mol-K – p^o = 0.1 MPa



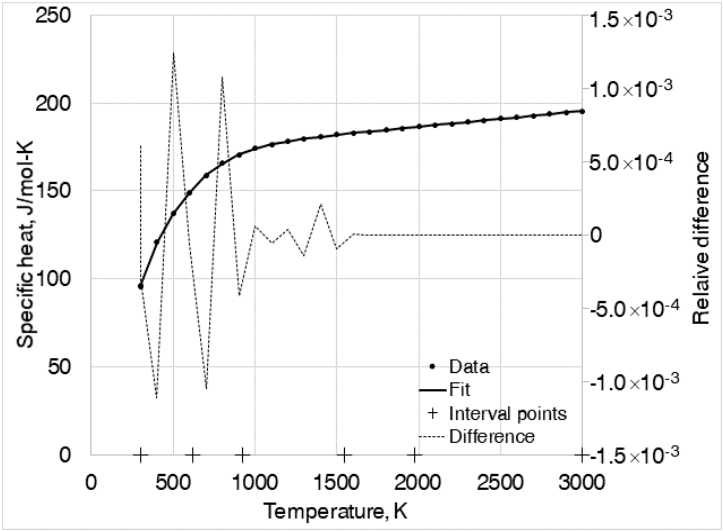


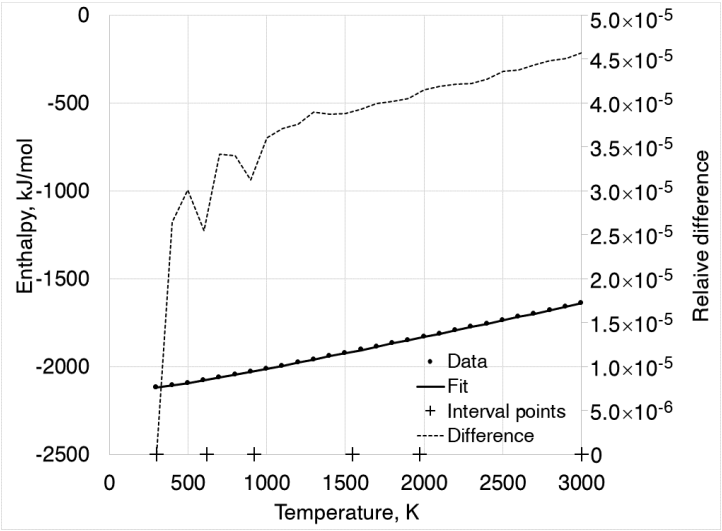


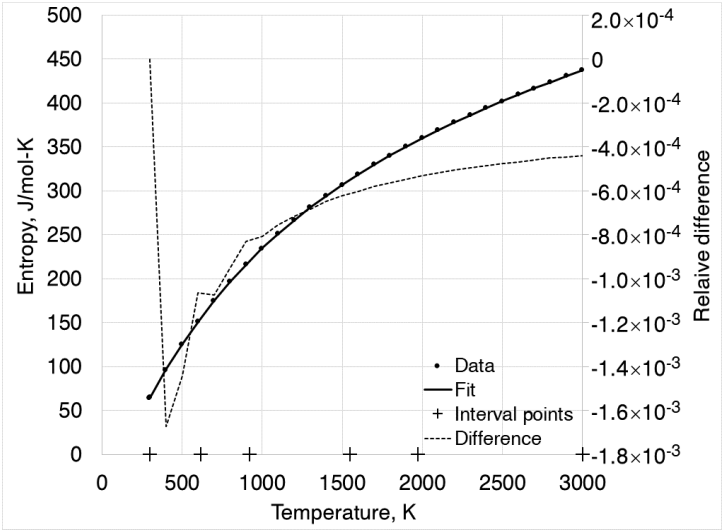


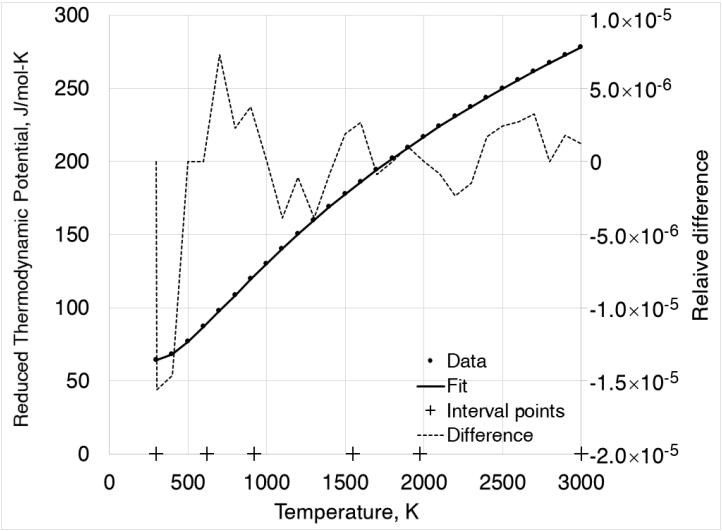
Beryllium Silicate (Be_2SiO_4) – Be2O4Si1(cr)

MW = 110.1075 g/mol – $\Delta_f H^o$(298.15 K) = -2117.225 kJ/mol – S^o(298.15 K) = 64.19 J/mol-K – p^o = 0.1 MPa



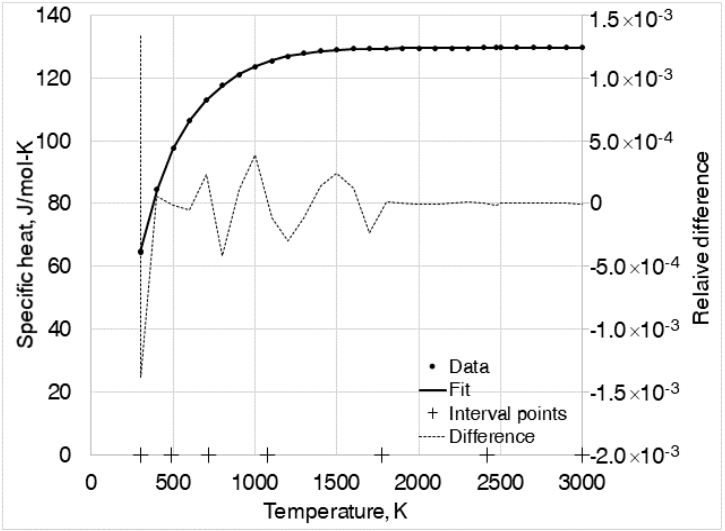


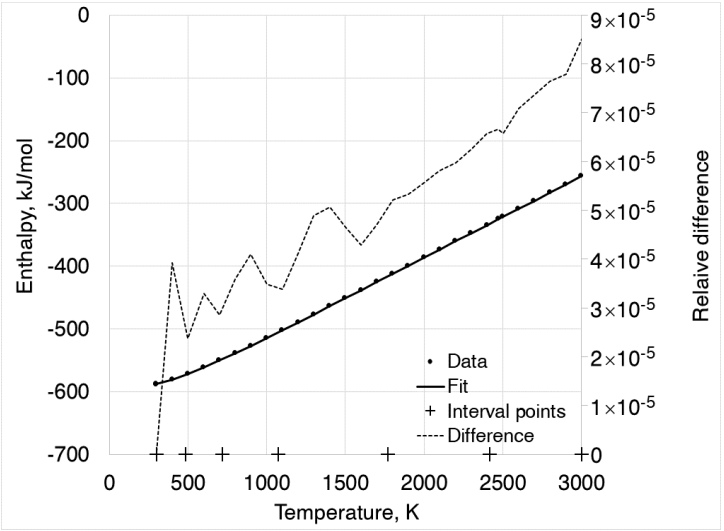


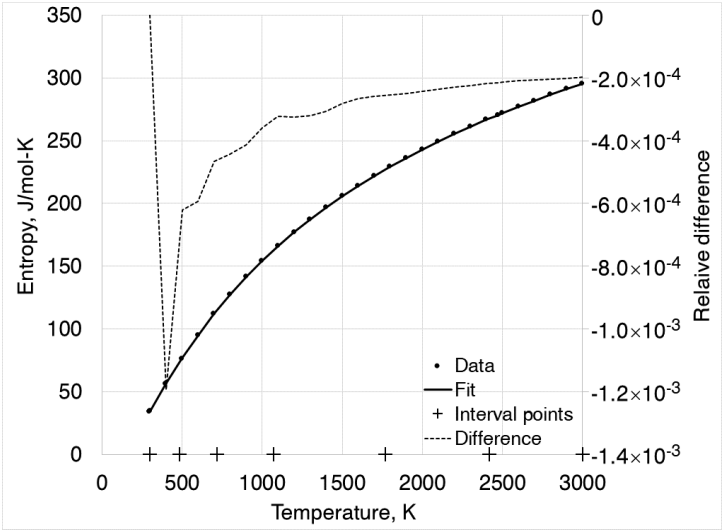


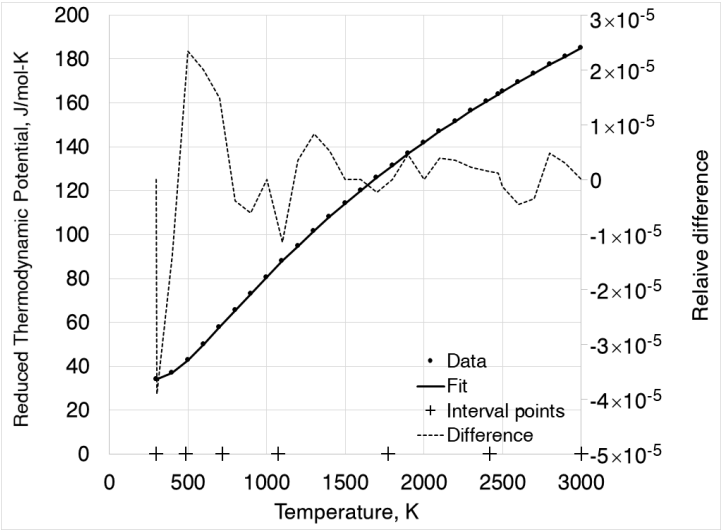
Beryllium Nitride, Alpha (Be_3N_2) – Be3N2(cr)

MW = 55.0499 g/mol – $\Delta_f H^o$(298.15 K) = -588.27 kJ/mol – S^o(298.15 K) = 34.129 J/mol-K – p^o = 0.1 MPa



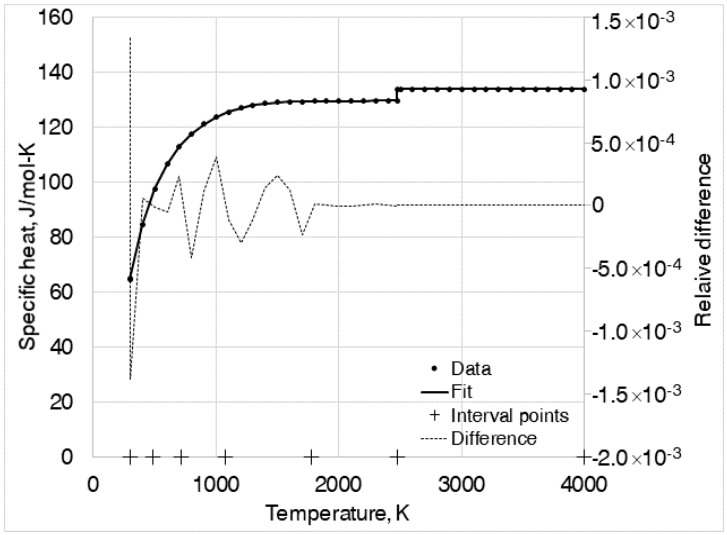


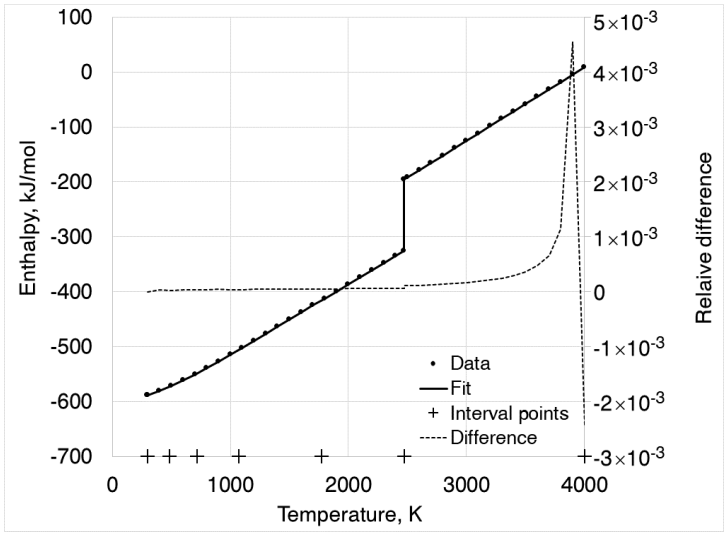


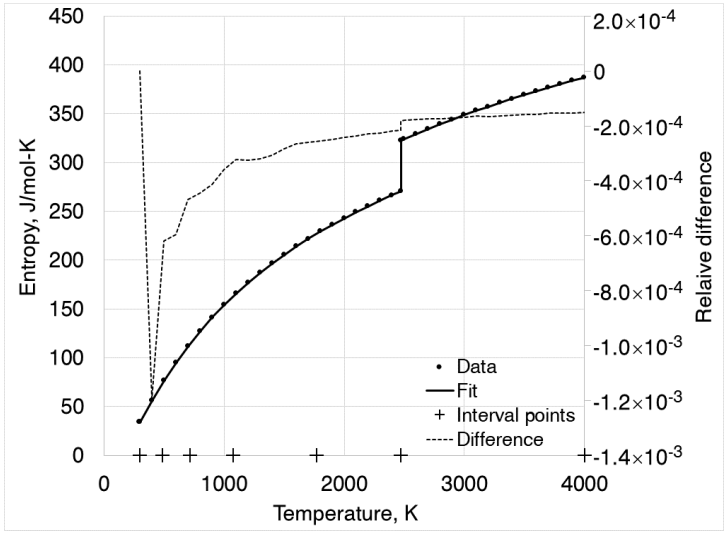


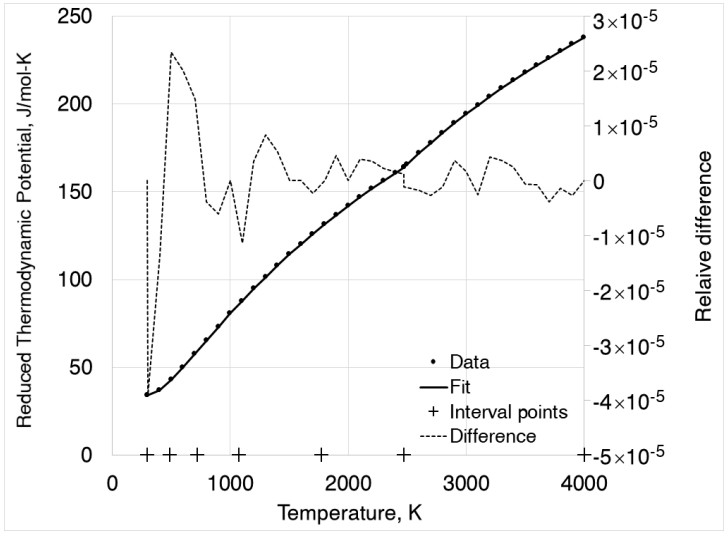
Beryllium Nitride (Be_3N_2) – Be3N2(cr,l)

MW = 55.0499 g/mol – $\Delta_f H^o$(298.15 K) = -588.27 kJ/mol – S^o(298.15 K) = 34.129 J/mol-K – p^o = 0.1 MPa



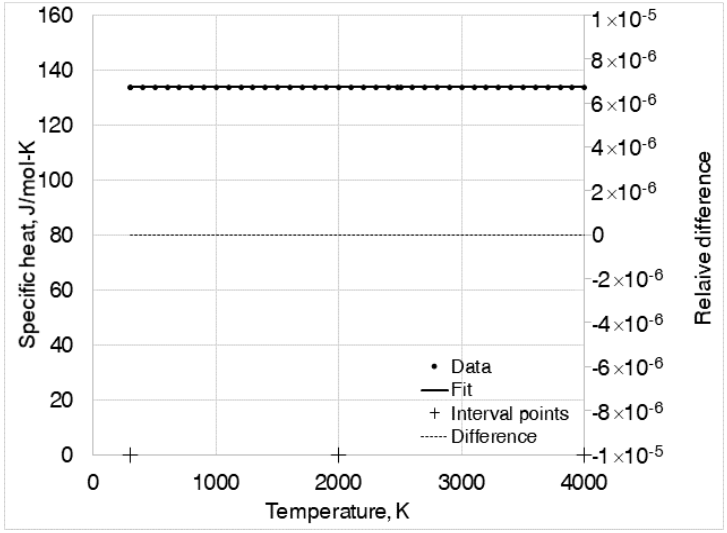


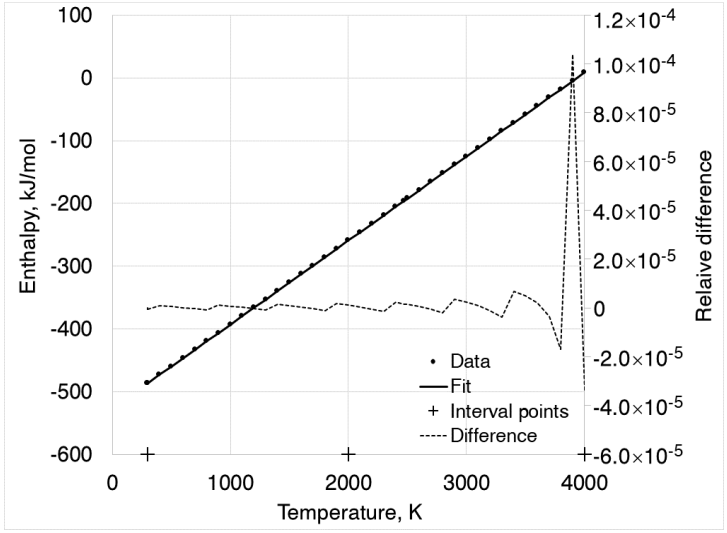


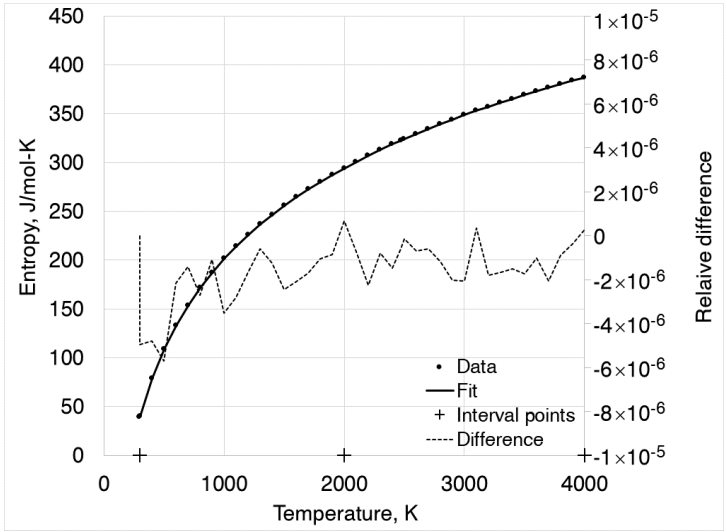


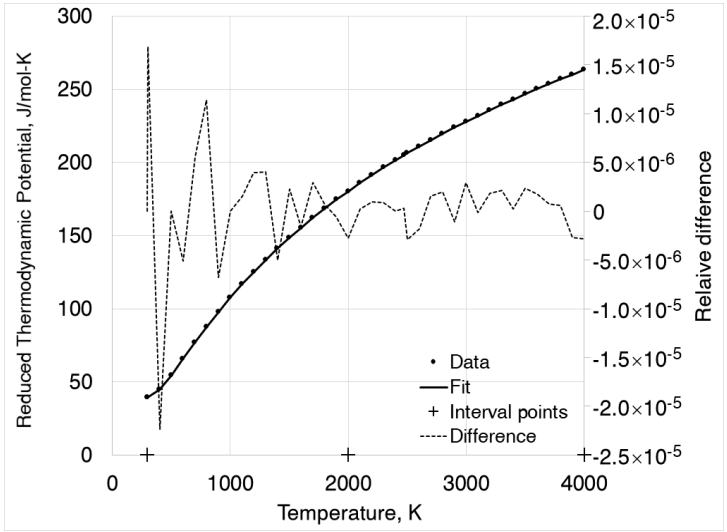
Beryllium Nitride (Be_3N_2) – Be3N2(l)

MW = 55.0499 g/mol – $\Delta_f H^o$(298.15 K) = -486.979 kJ/mol – S^o(298.15 K) = 39.457 J/mol-K – p^o = 0.1 MPa



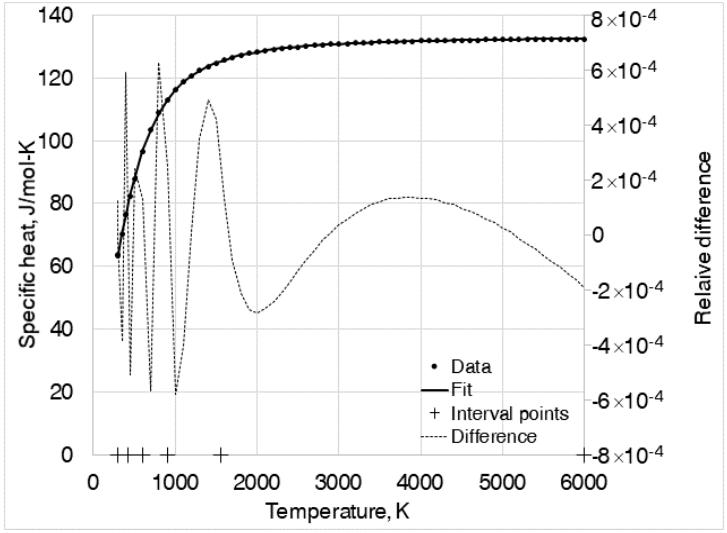


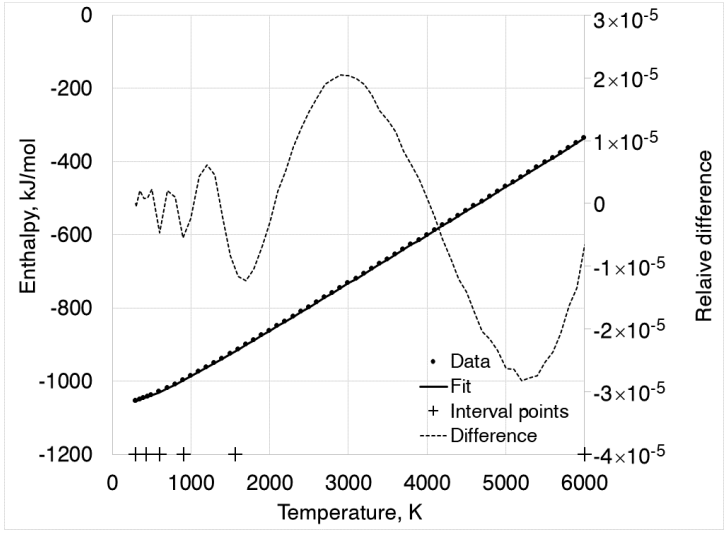


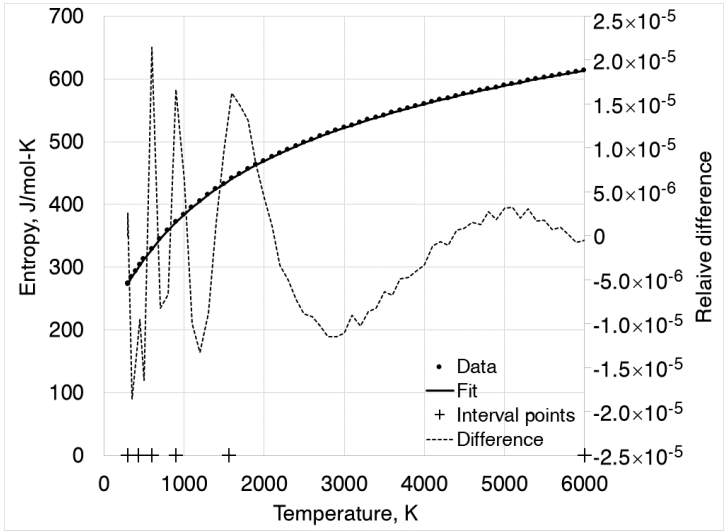


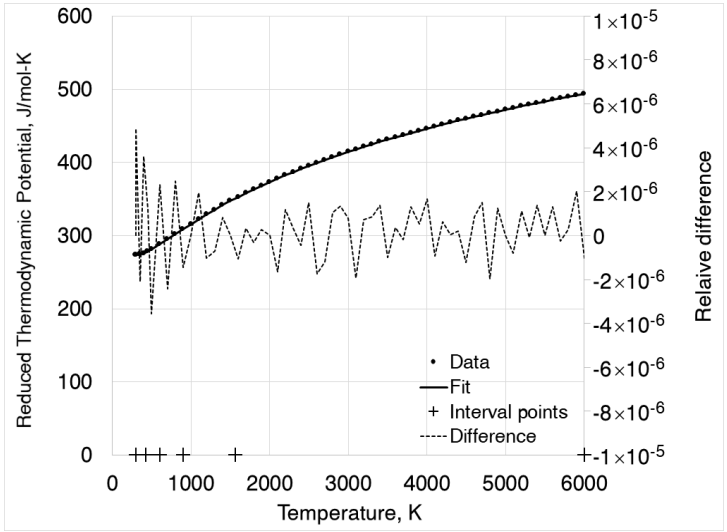
Beryllium Oxide ($(BeO)_3$) – Be3O3(g)

MW = 75.0347 g/mol – $\Delta_f H^o$(298.15 K) = -1054.368 kJ/mol – S^o(298.15 K) = 273.337 J/mol-K – p^o = 0.1 MPa



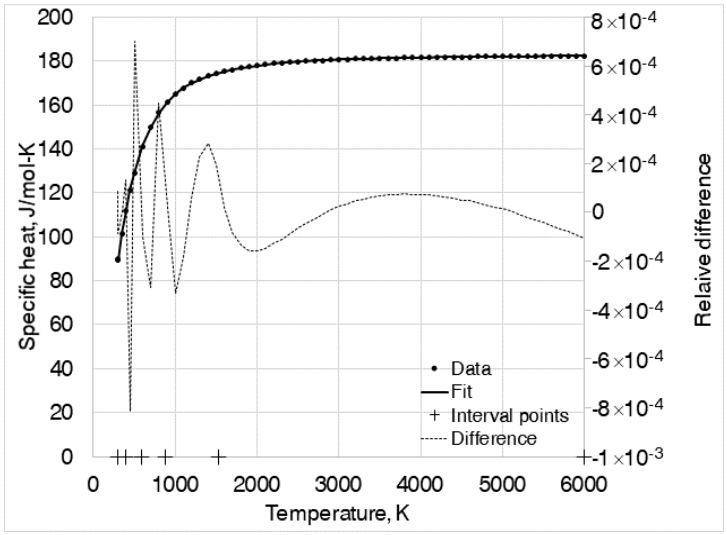


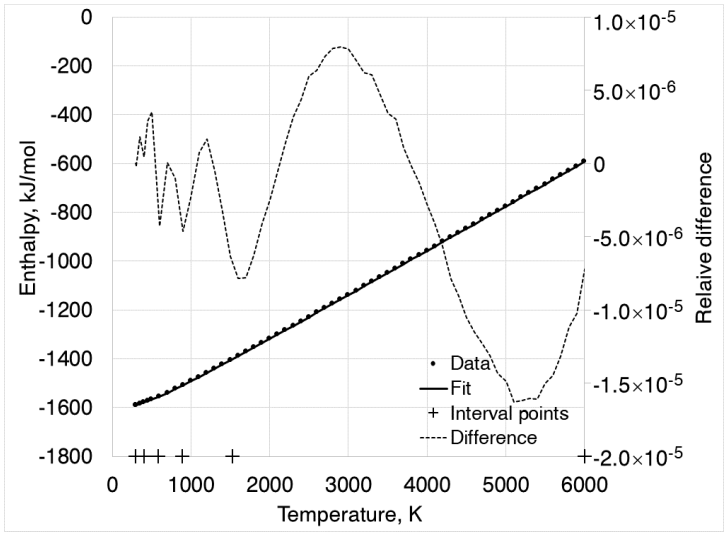


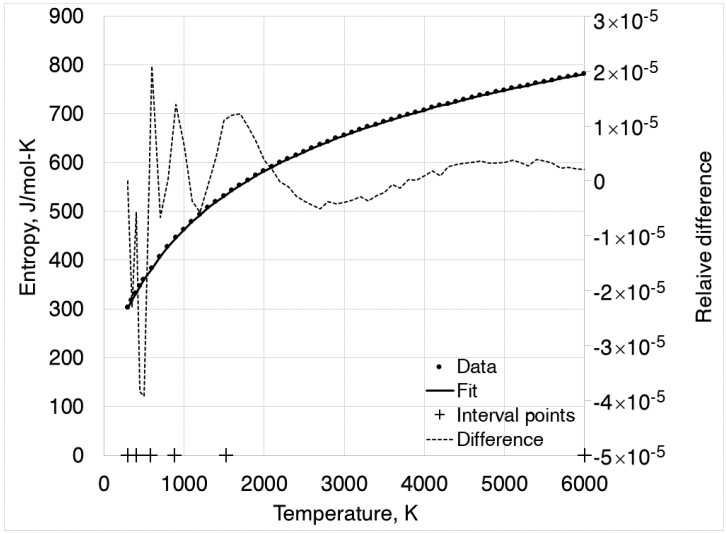


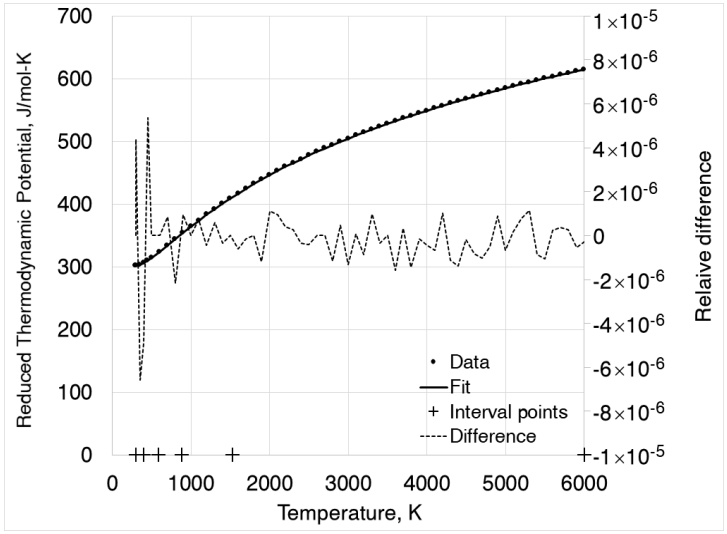
Beryllium Oxide ((BeO)$_4$) – Be4O4(g)

MW = 100.0463 g/mol – $\Delta_f H^o$(298.15 K) = -1589.92 kJ/mol – S^o(298.15 K) = 302.575 J/mol-K – p^o = 0.1 MPa



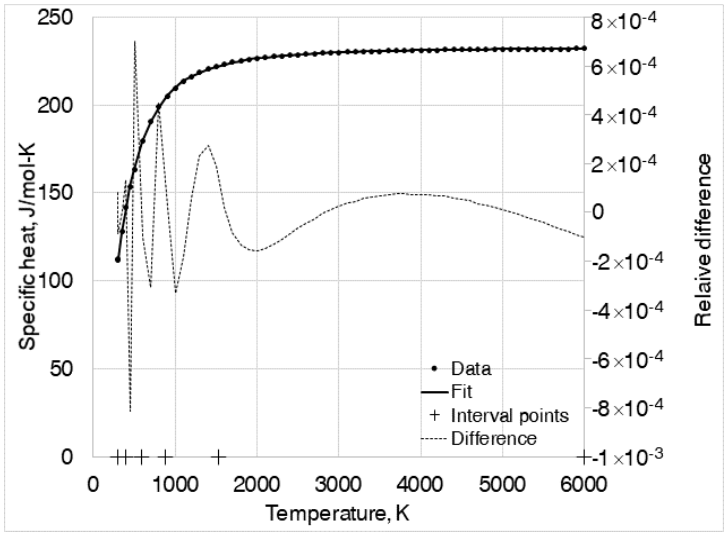


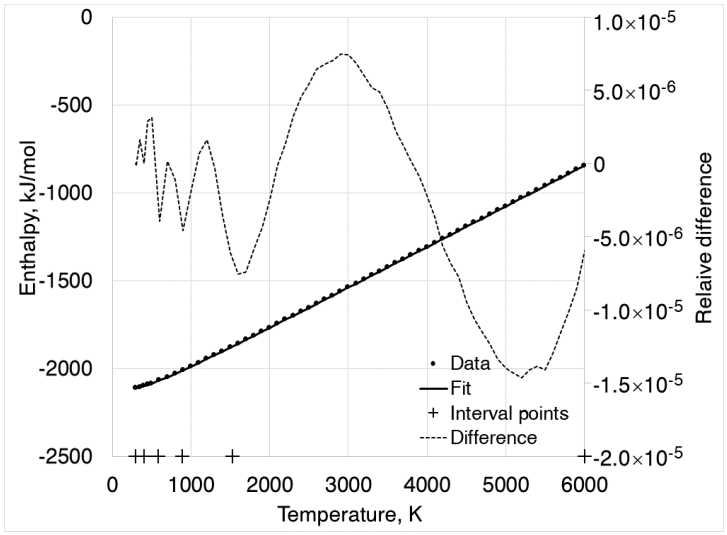


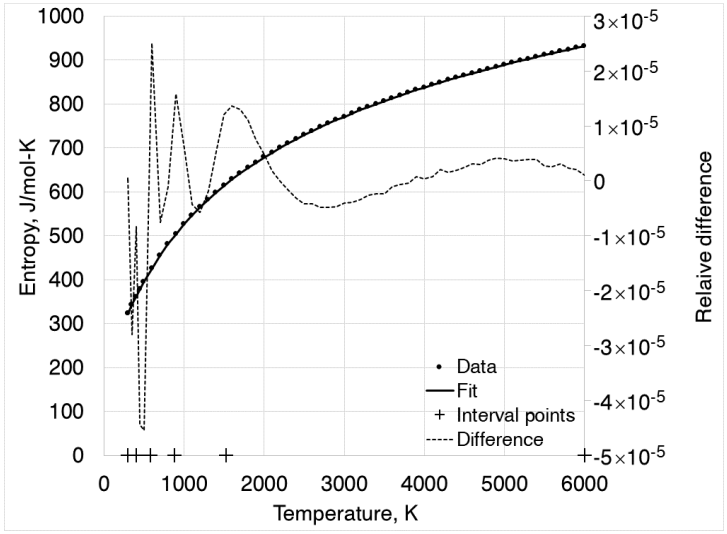


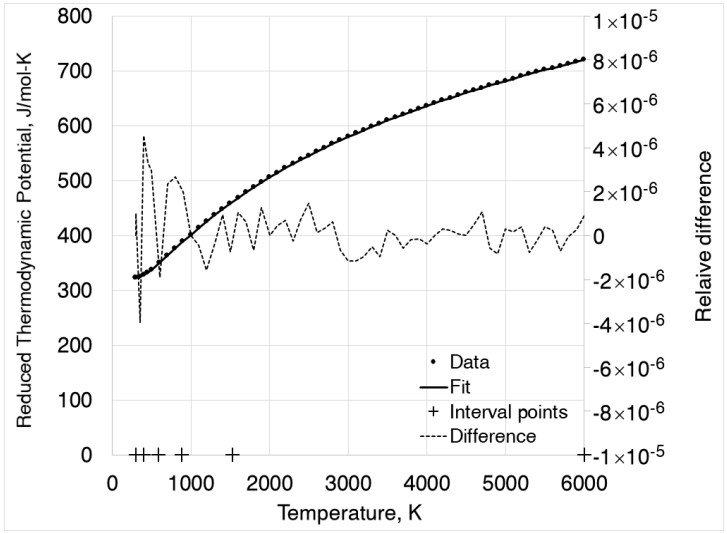
Beryllium Oxide ($(BeO)_5$) – Be5O5(g)

MW = 125.0579 g/mol – $\Delta_f H^o$(298.15 K) = -2112.92 kJ/mol – S^o(298.15 K) = 323.278 J/mol-K – p^o = 0.1 MPa



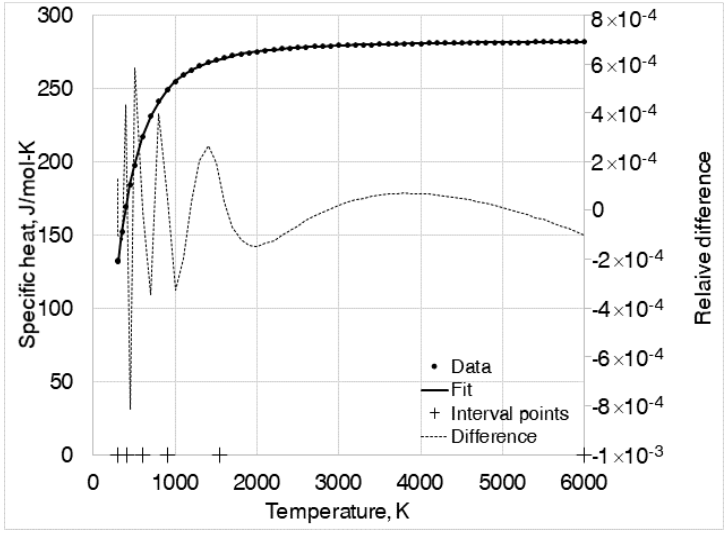


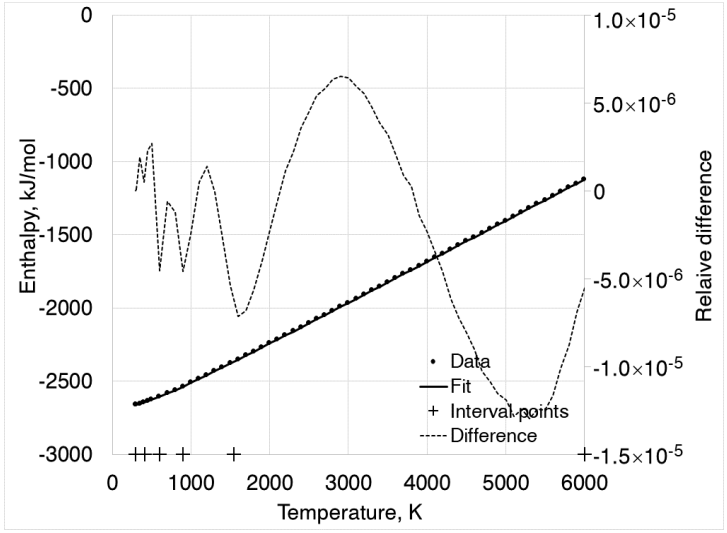


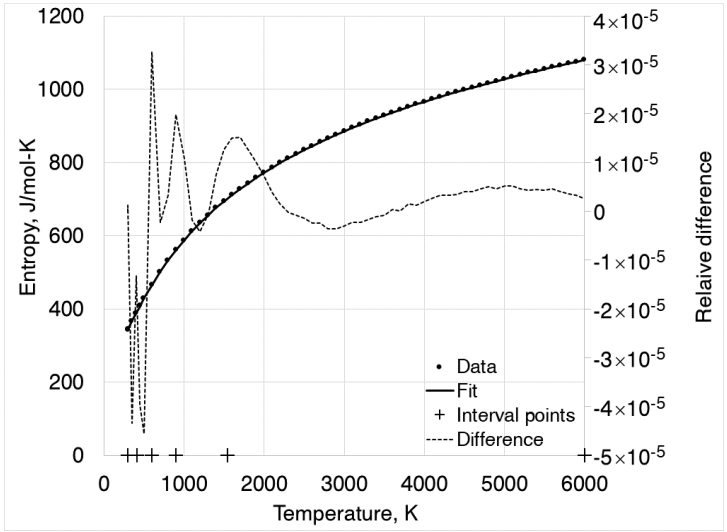


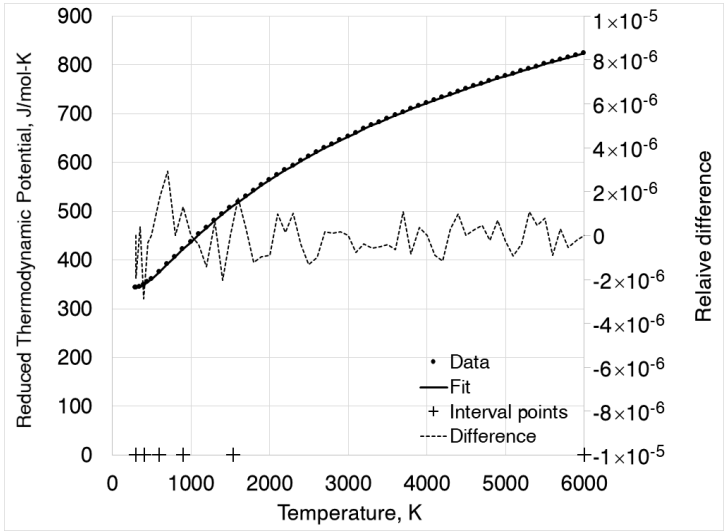
Beryllium Oxide ($(BeO)_6$) – Be6O6(g)

MW = 150.0695 g/mol – $\Delta_f H^o$(298.15 K) = -2661.024 kJ/mol – S^o(298.15 K) = 343.191 J/mol-K – p^o = 0.1 MPa



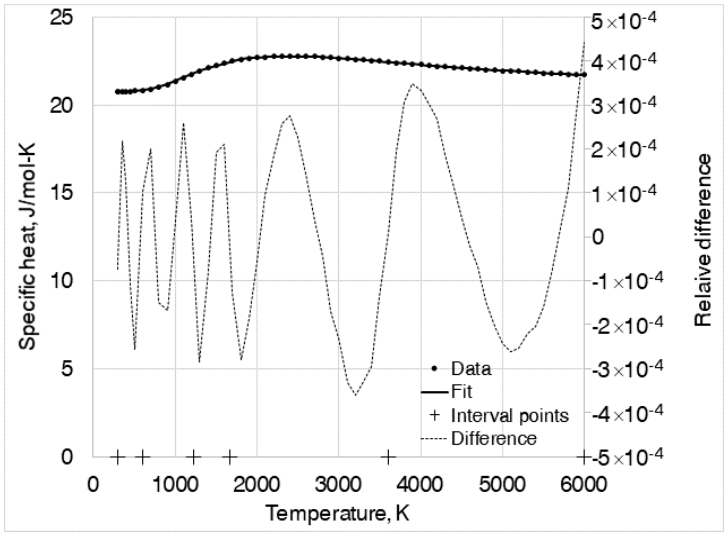


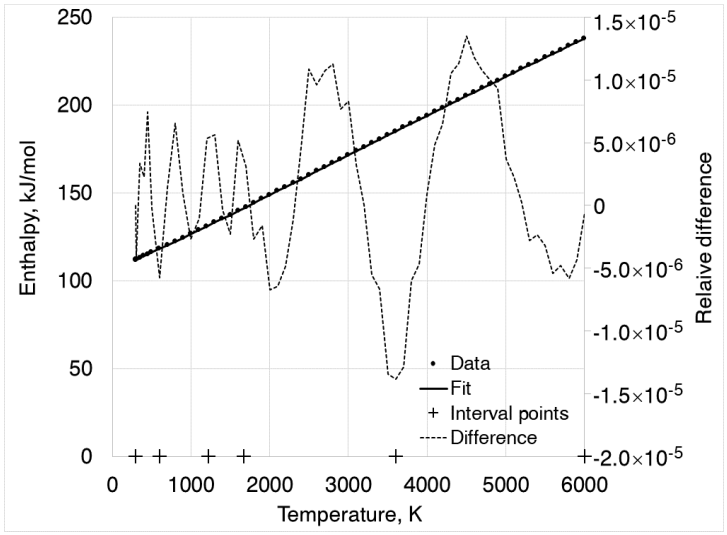


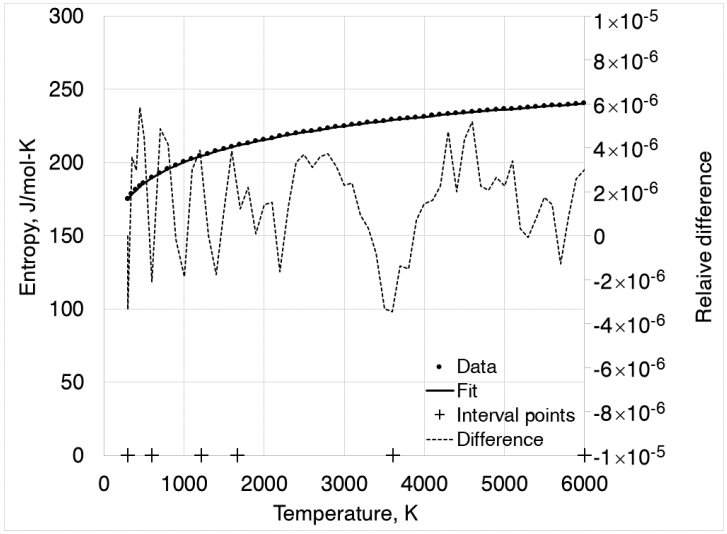


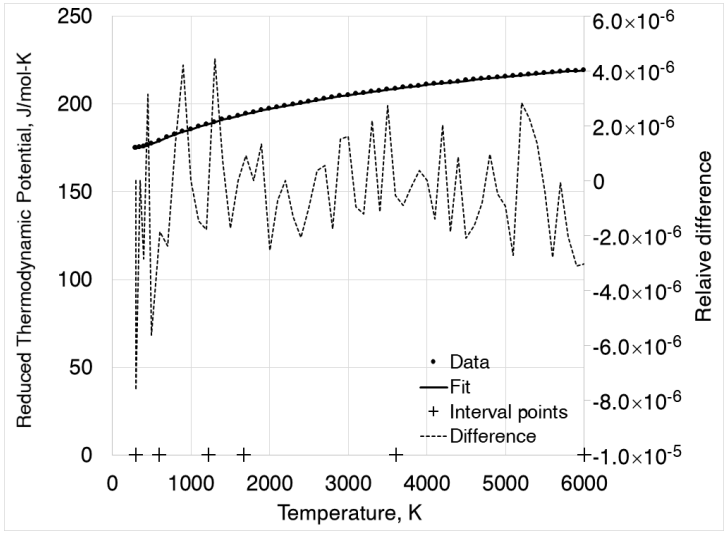
Bromine (Br) – Br1(g)

MW = 79.904 g/mol – $\Delta_f H^o$(298.15 K) = 111.86 kJ/mol – S^o(298.15 K) = 175.017 J/mol-K – p^o = 0.1 MPa



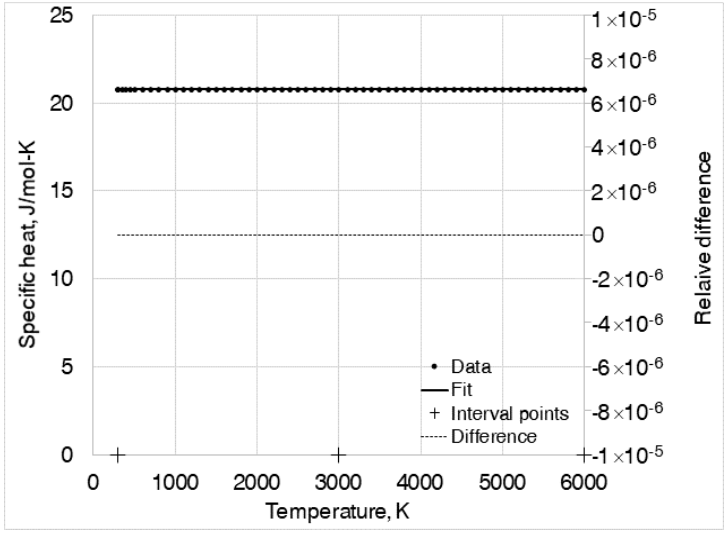


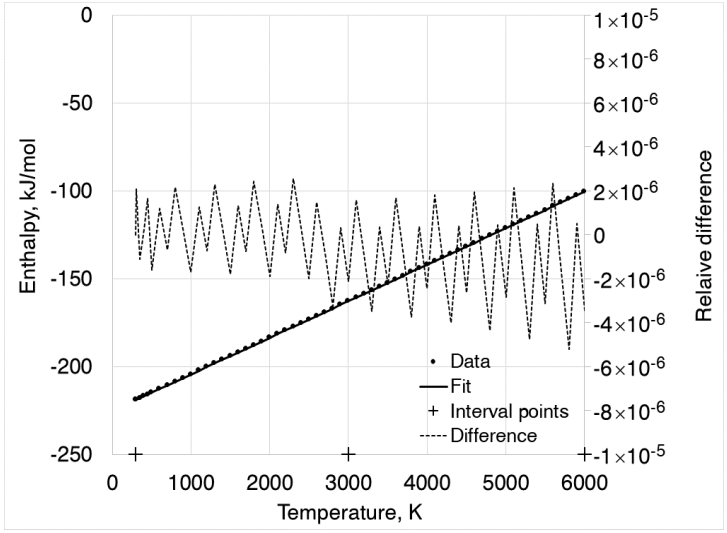


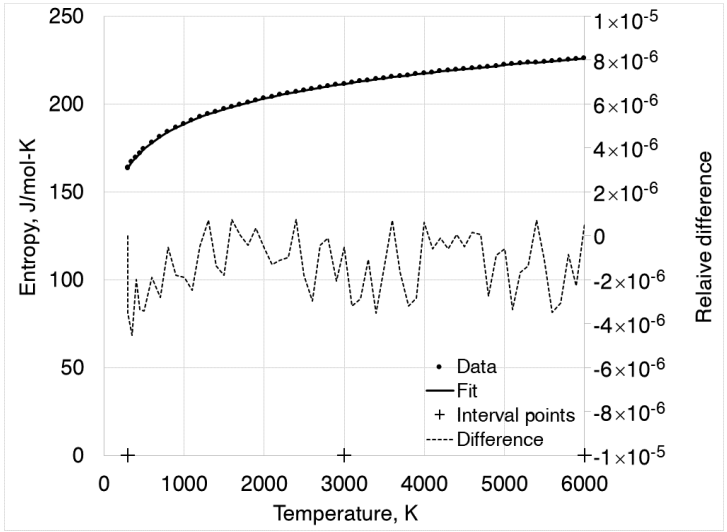


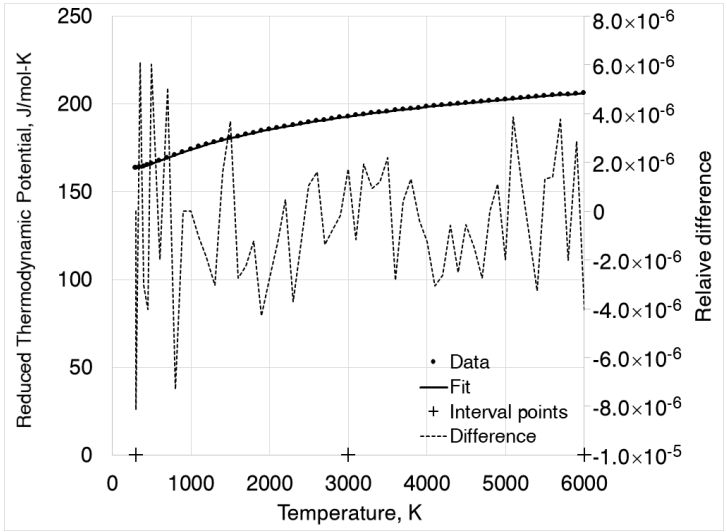
Bromine, Ion (Br^-) – Br1-(g)

MW = 79.905 g/mol – $\Delta_f H^o$(298.15 K) = -219.008 kJ/mol – S^o(298.15 K) = 163.491 J/mol-K – p^o = 0.1 MPa



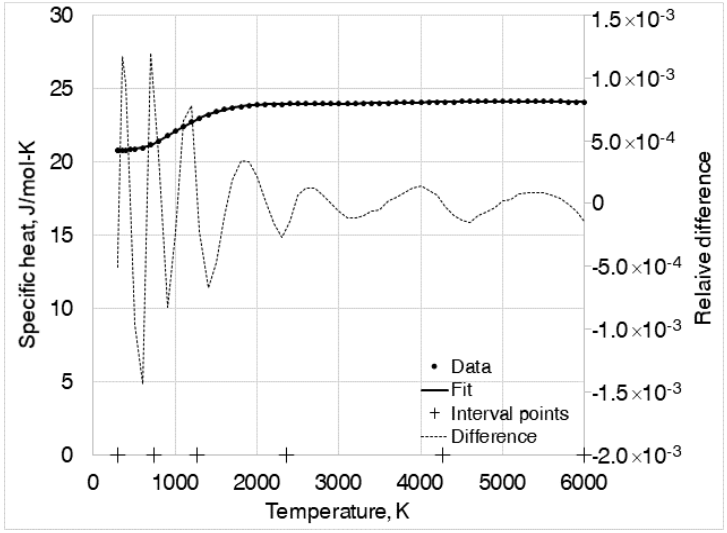


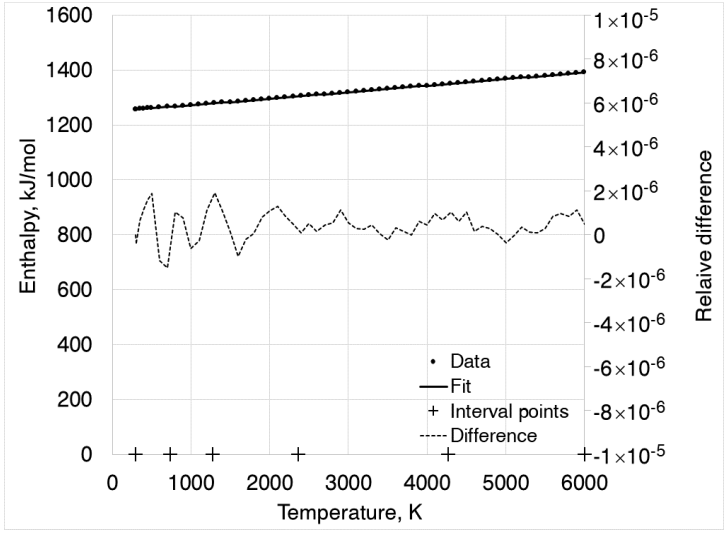


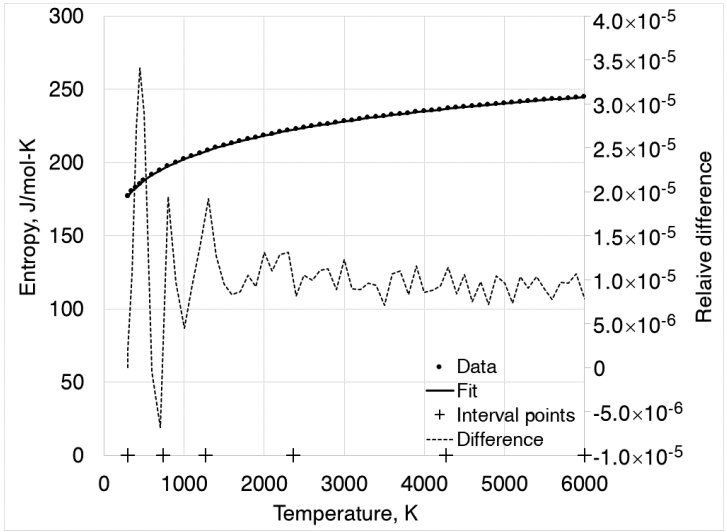


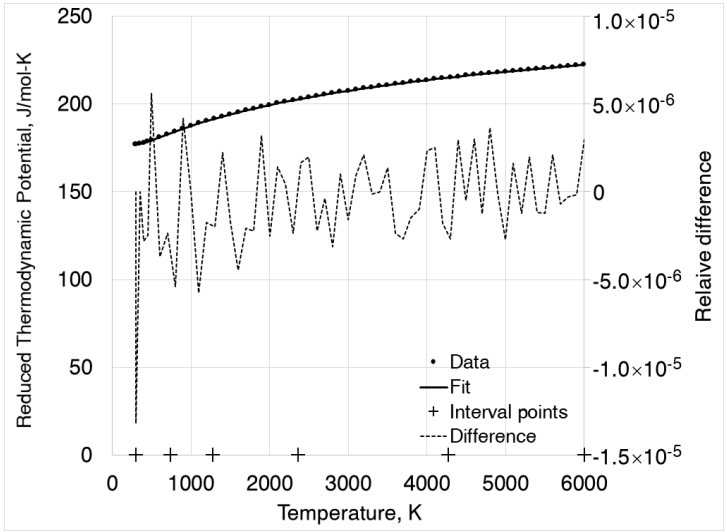
Bromine, Ion (Br^+) – Br1+(g)

MW = 79.903 g/mol – $\Delta_f H^o$(298.15 K) = 1257.091 kJ/mol – S^o(298.15 K) = 176.872 J/mol-K – p^o = 0.1 MPa



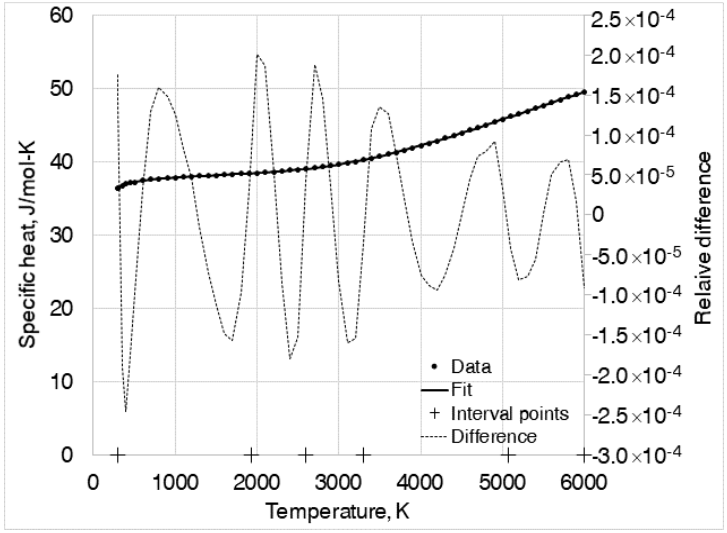


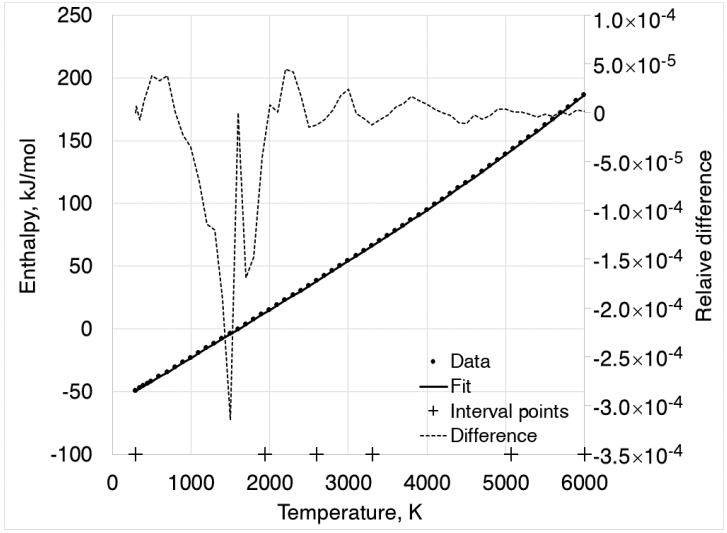


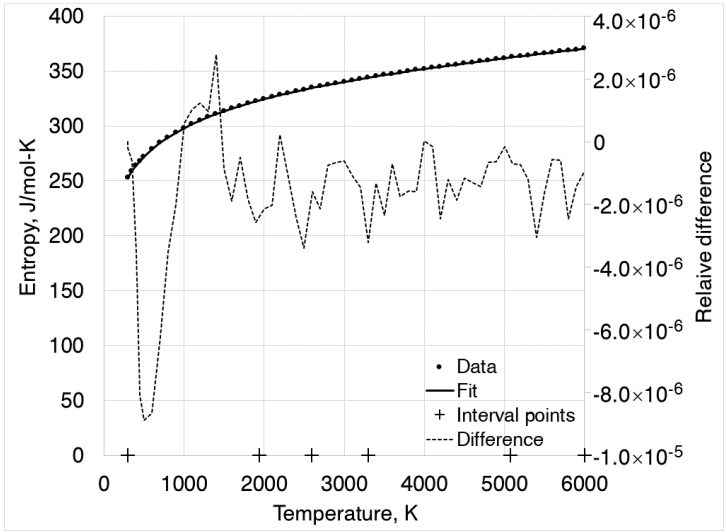


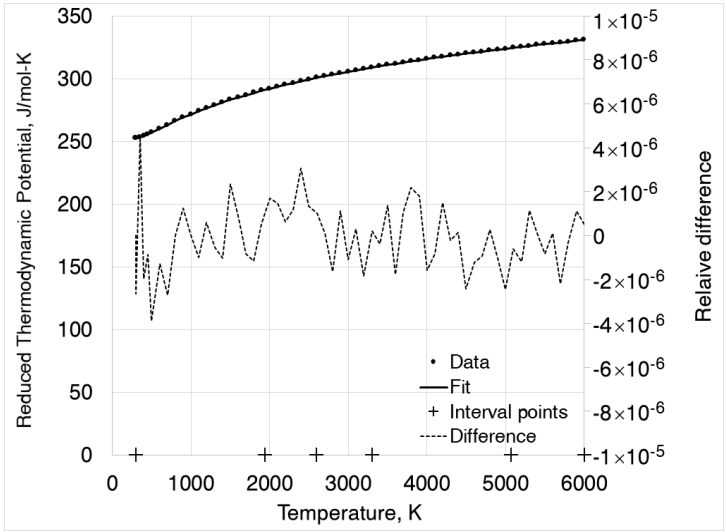
Calcium Bromide (CaBr) – Br1Ca1(g)

MW = 119.98 g/mol – $\Delta_f H^o$(298.15 K) = -49.396 kJ/mol – S^o(298.15 K) = 252.894 J/mol-K – p^o = 0.1 MPa



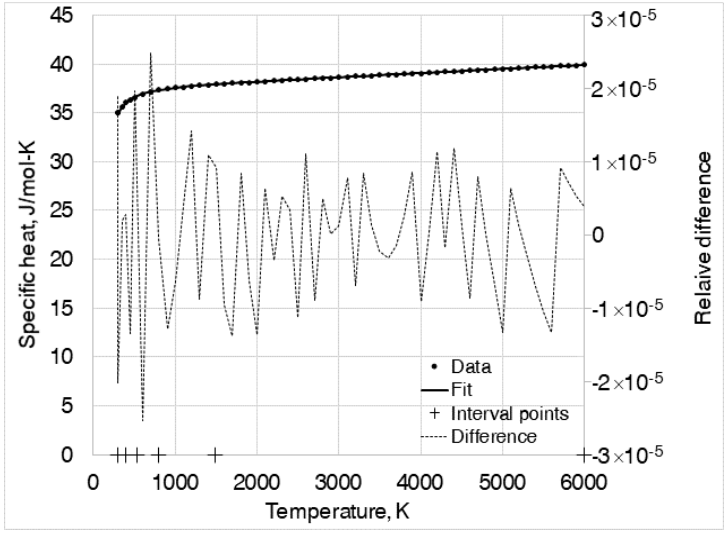


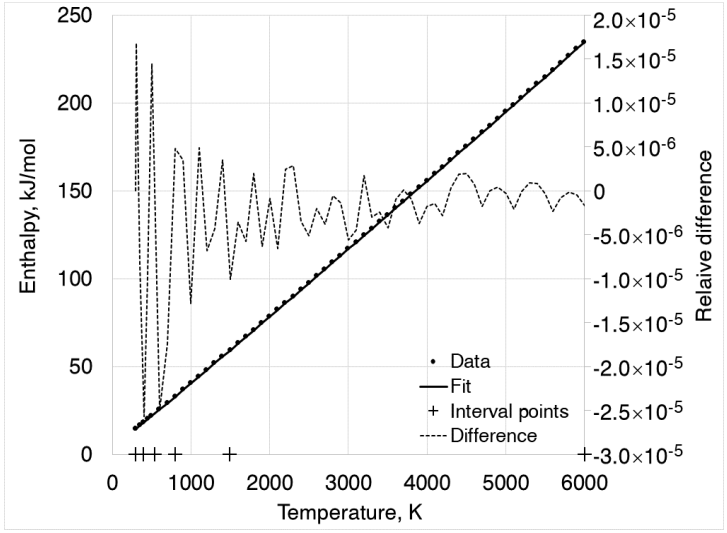


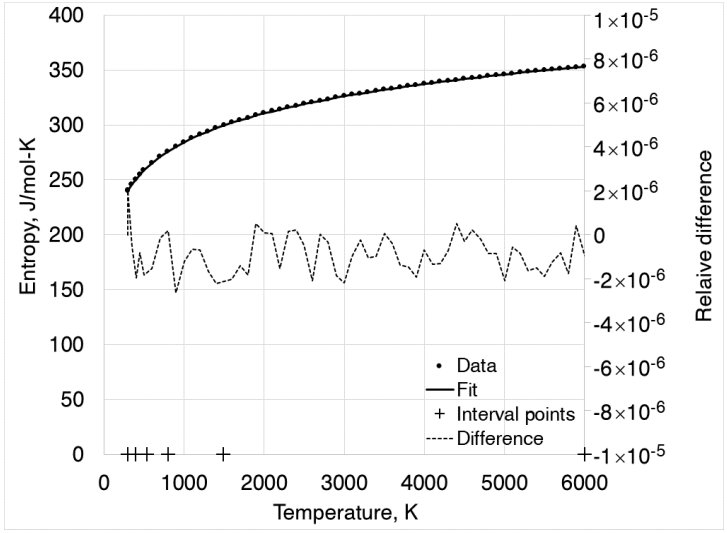


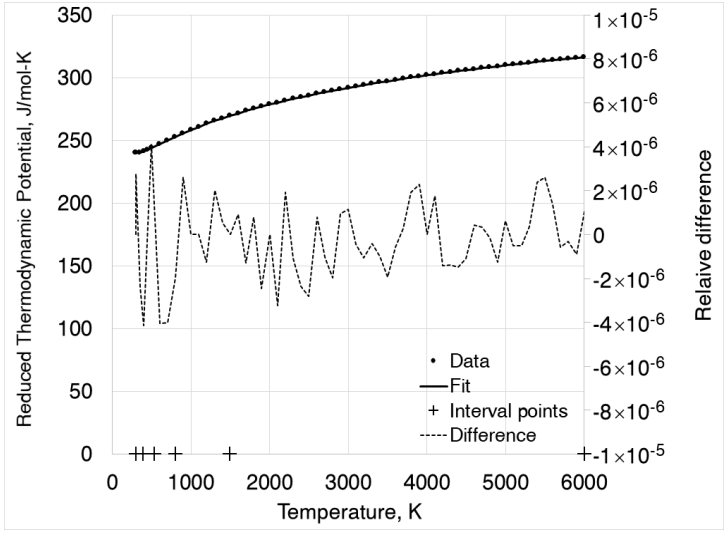
Bromine Chloride (BrCl) – Br1Cl1(g)

MW = 115.357 g/mol – $\Delta_f H^o$(298.15 K) = 14.644 kJ/mol – S^o(298.15 K) = 240.001 J/mol-K – p^o = 0.1 MPa



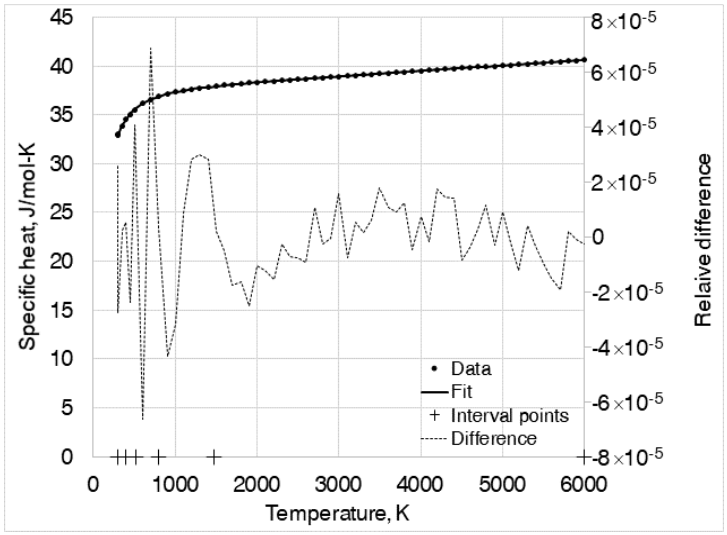


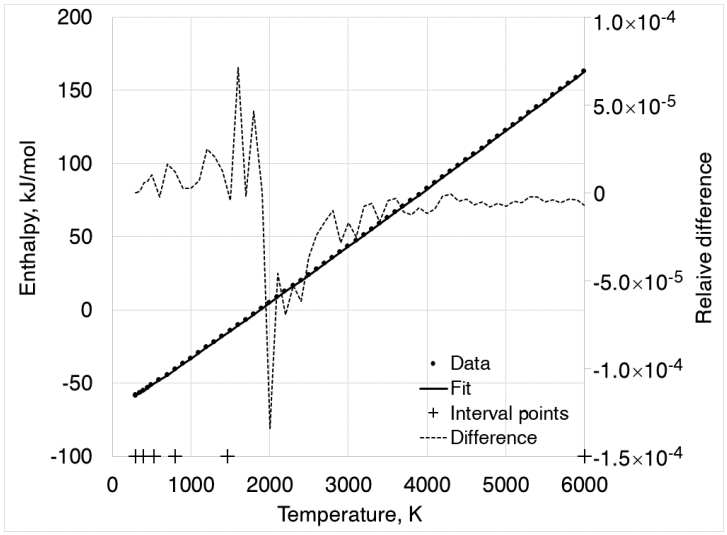


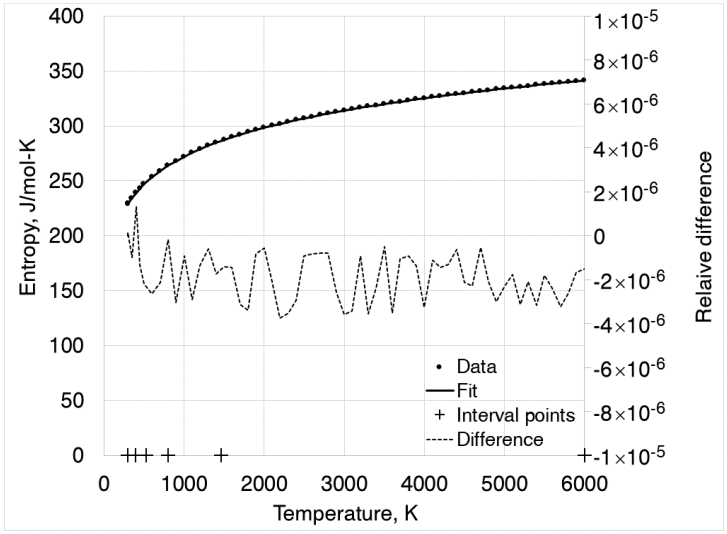


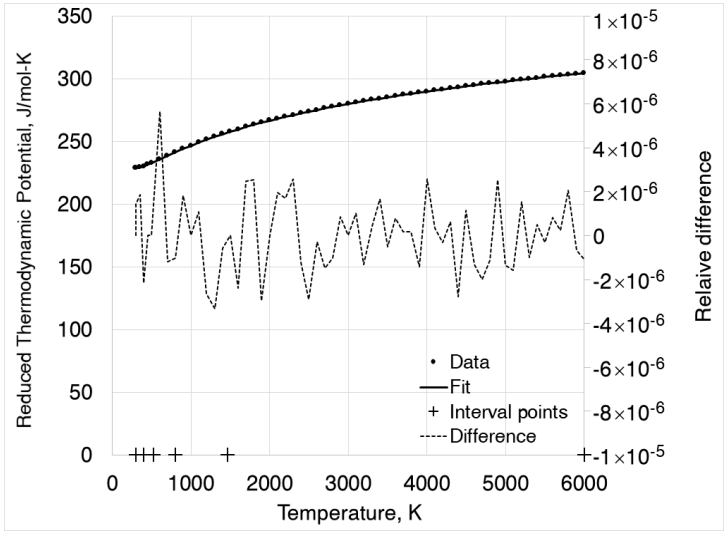
Bromine Fluoride (BrF) – Br1F1(g)

MW = 98.902 g/mol – $\Delta_f H^o$(298.15 K) = -58.463 kJ/mol – S^o(298.15 K) = 228.967 J/mol-K – p^o = 0.1 MPa



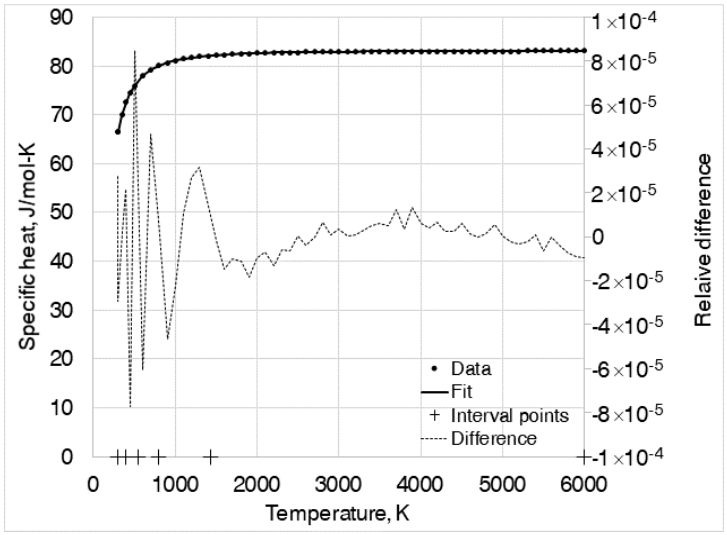


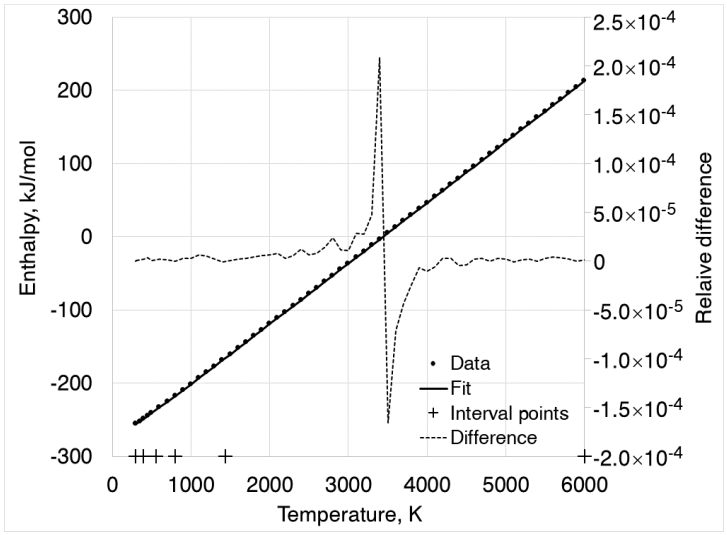


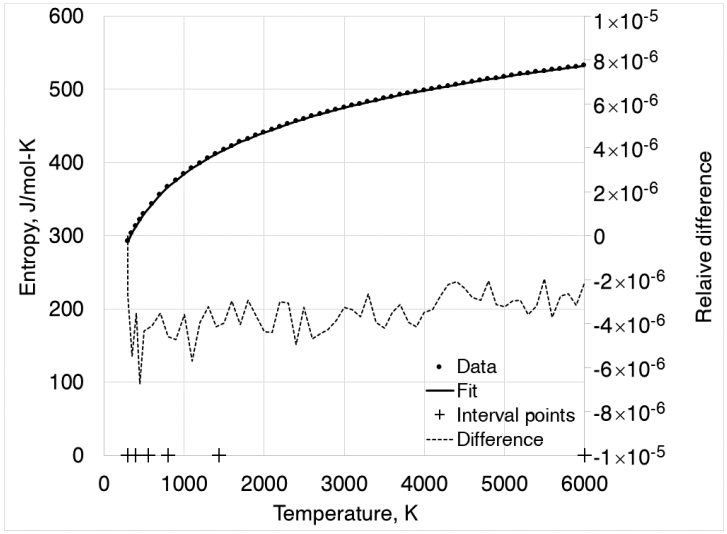


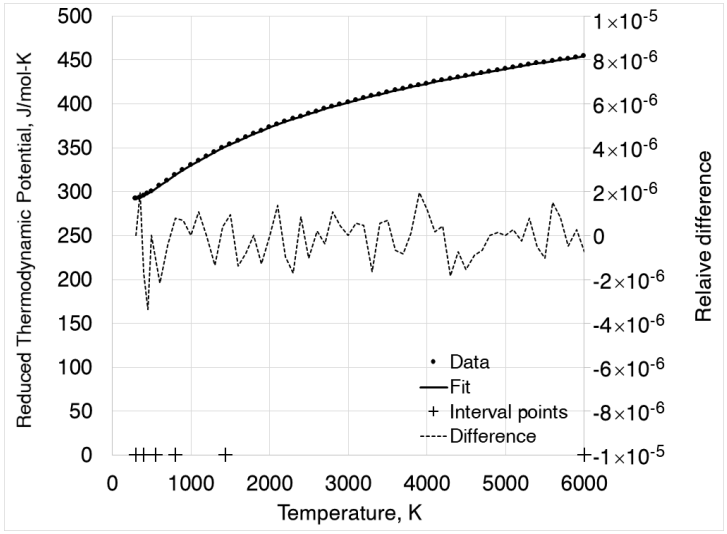
Bromine Fluoride (BrF_3) – Br1F3(g)

MW = 136.899 g/mol – $\Delta_f H^o$(298.15 K) = -255.588 kJ/mol – S^o(298.15 K) = 292.397 J/mol-K – p^o = 0.1 MPa



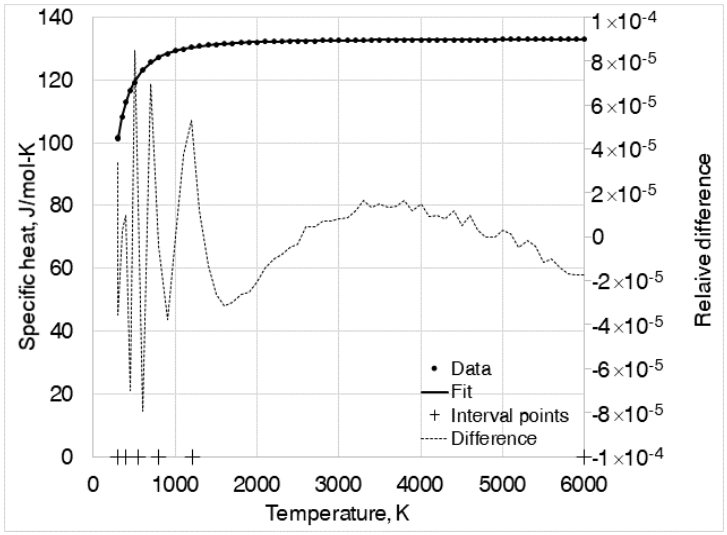


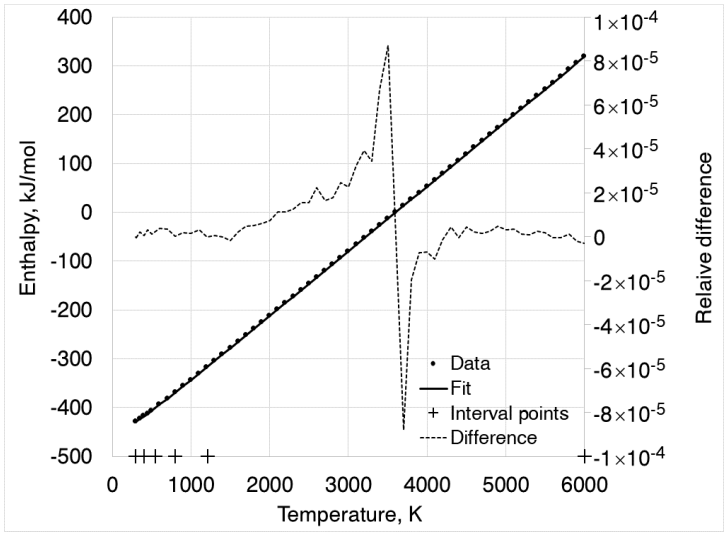


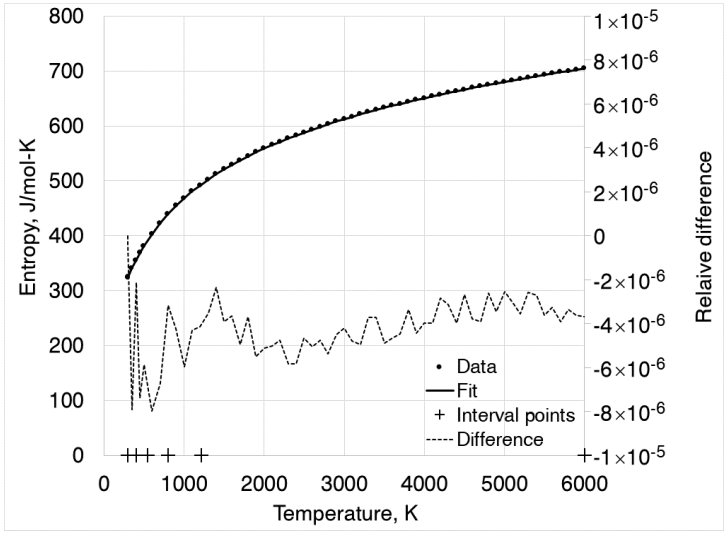


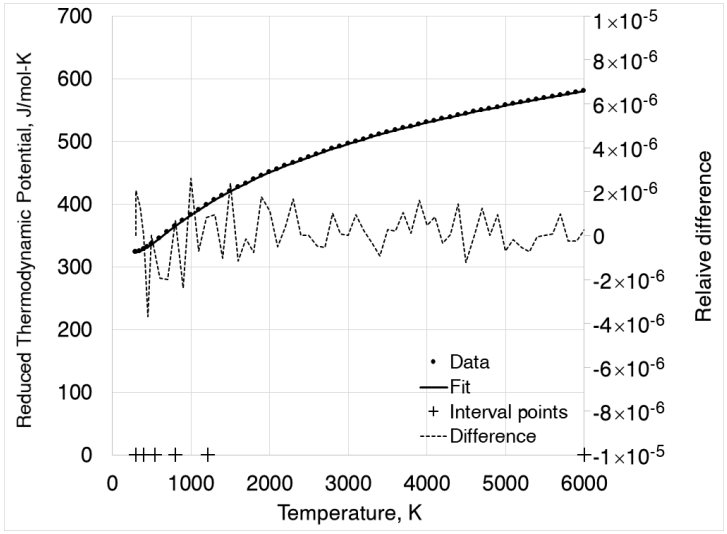
Bromine Fluoride (BrF_5) – Br1F5(g)

MW = 174.896 g/mol – $\Delta_f H^o$(298.15 K) = -428.718 kJ/mol – S^o(298.15 K) = 323.682 J/mol-K – p^o = 0.1 MPa



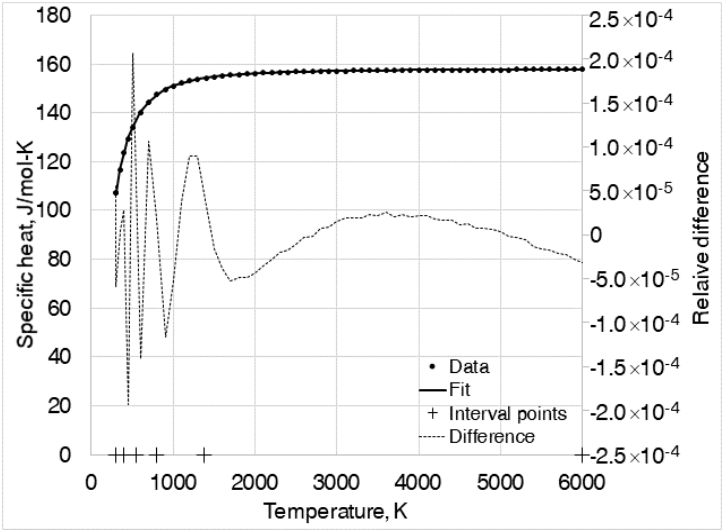


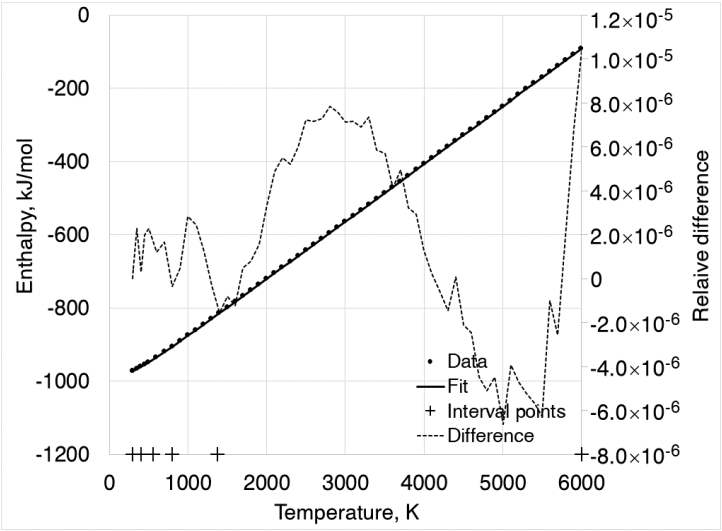


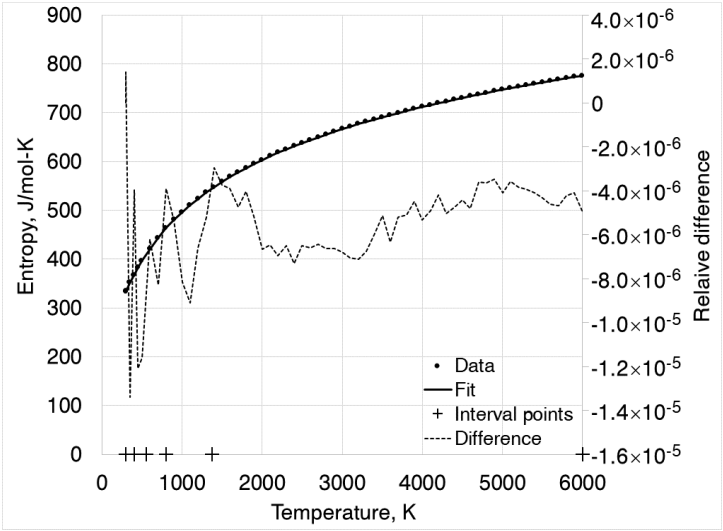


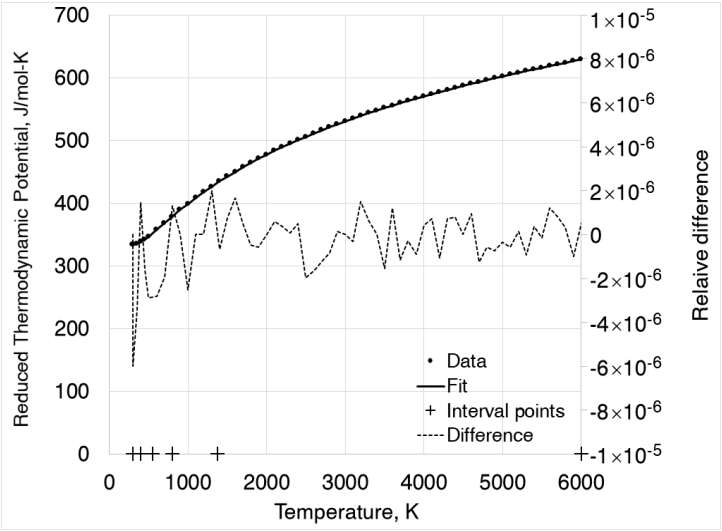
Sulfur Bromide Fluoride ($SBrF_5$) – Br1F5S1(g)

MW = 206.96 g/mol – $\Delta_f H^o$(298.15 K) = -972.822 kJ/mol – S^o(298.15 K) = 333.648 J/mol-K – p^o = 0.1 MPa



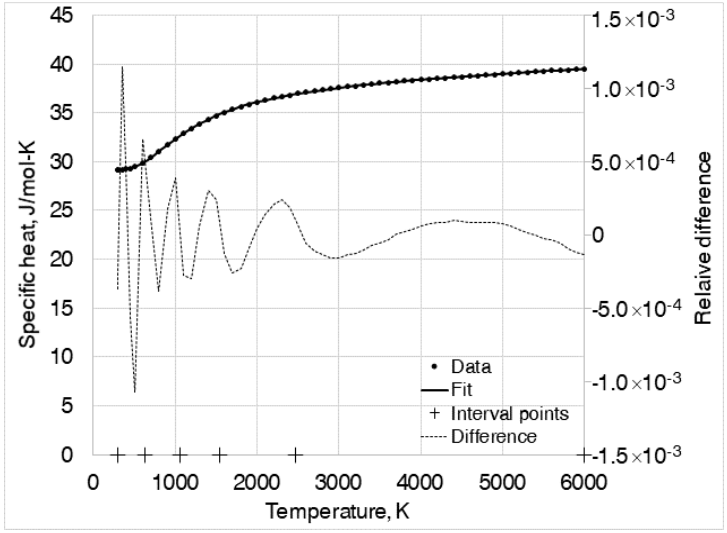


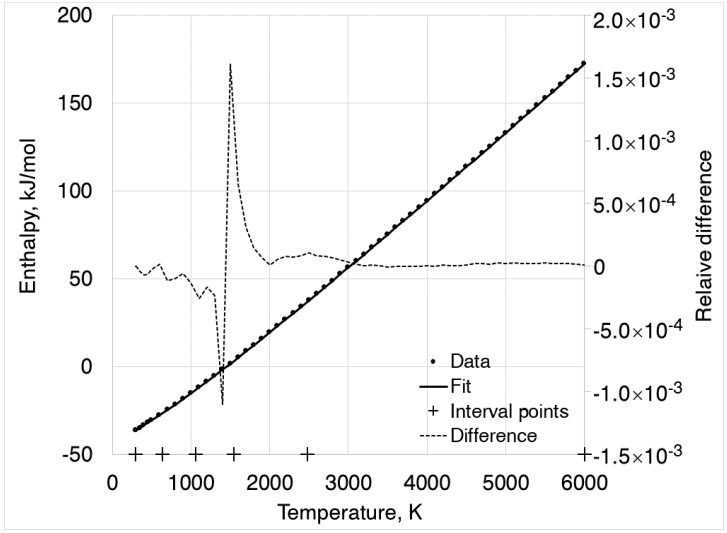


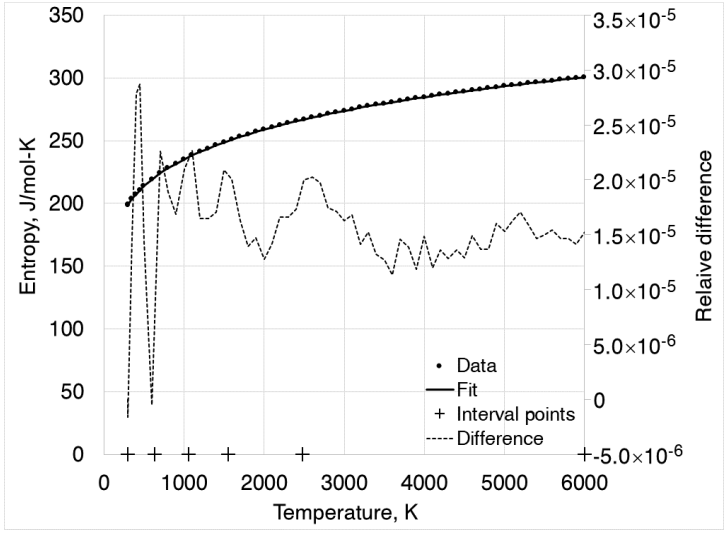


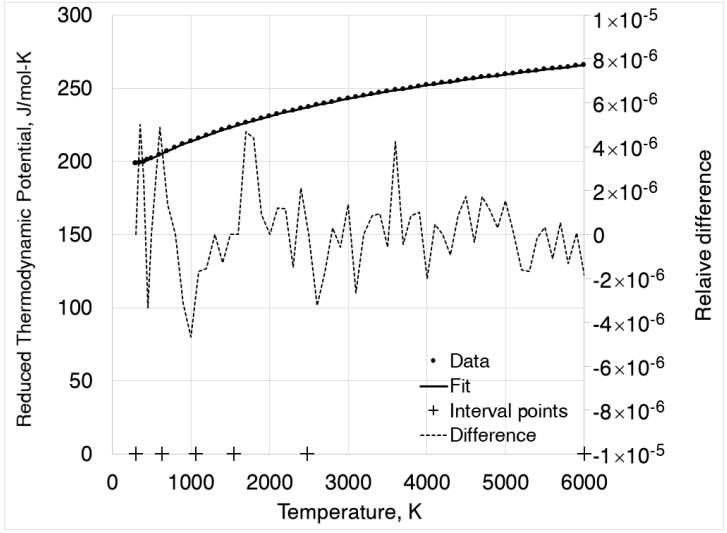
Hydrogen Bromide (HBr) – Br1H1(g)

MW = 80.912 g/mol – $\Delta_f H^o$(298.15 K) = -36.443 kJ/mol – S^o(298.15 K) = 198.699 J/mol-K – p^o = 0.1 MPa



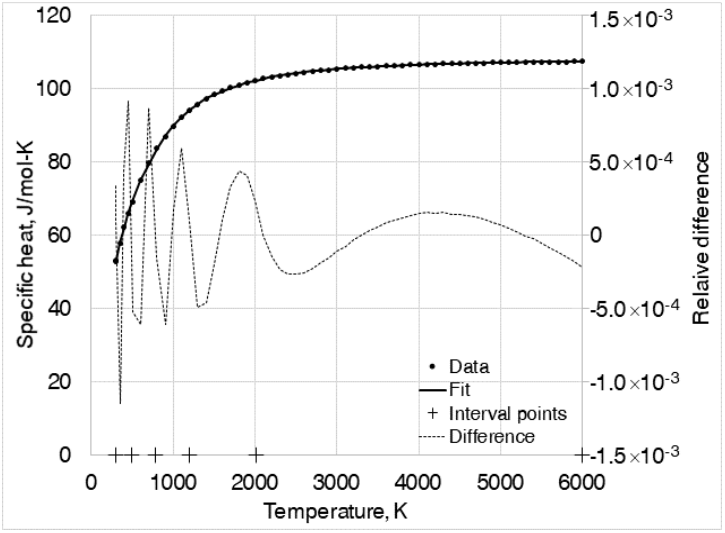


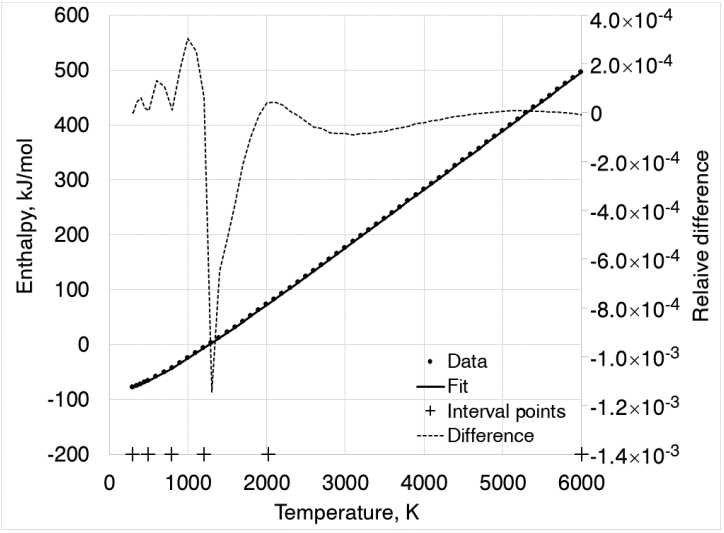


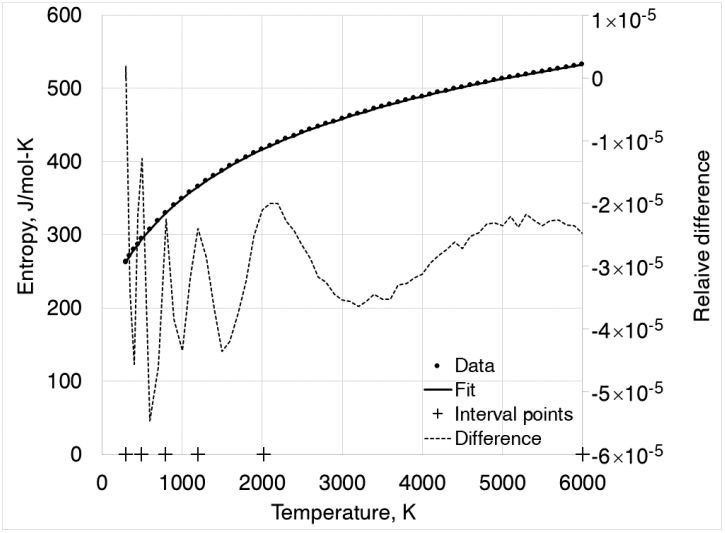


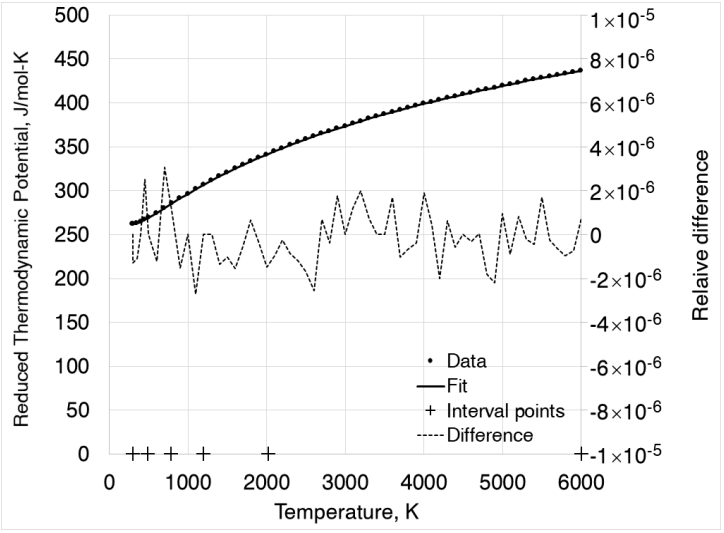
Bromosilane (SiH_3Br) – Br1H3Si1(g)

MW = 111.013 g/mol – $\Delta_f H^o$(298.15 K) = -78.241 kJ/mol – S^o(298.15 K) = 262.466 J/mol-K – p^o = 0.1 MPa



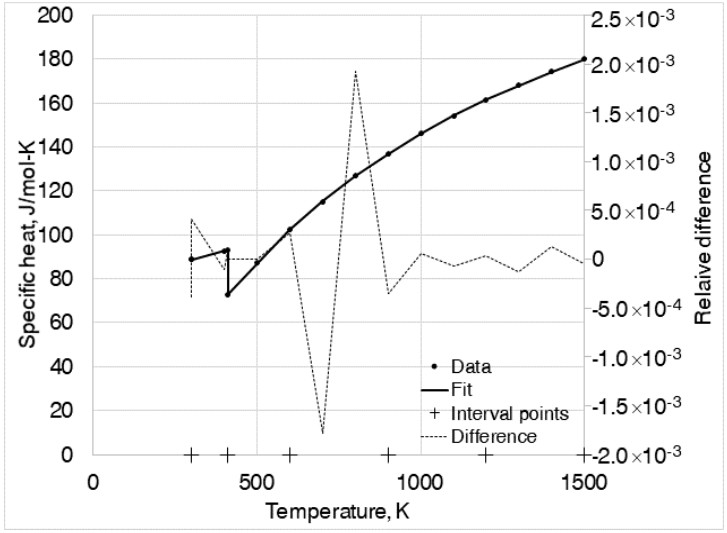


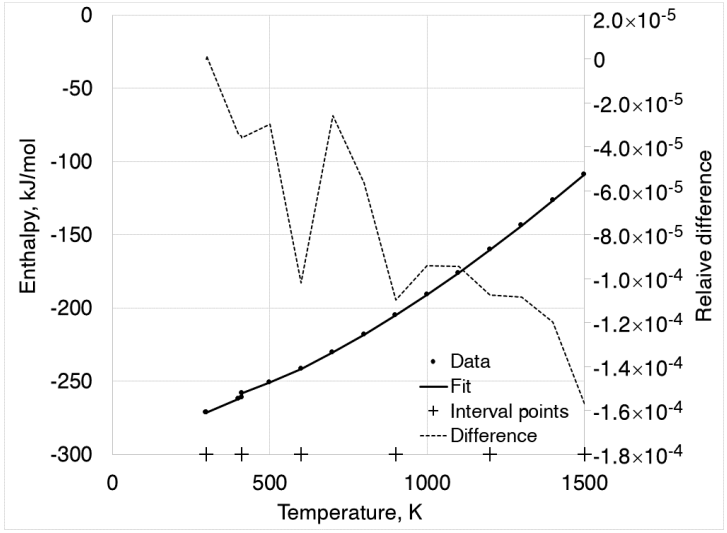


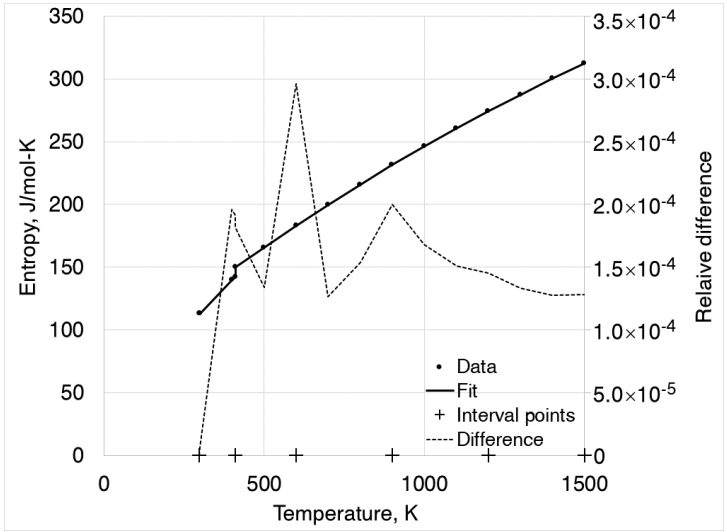


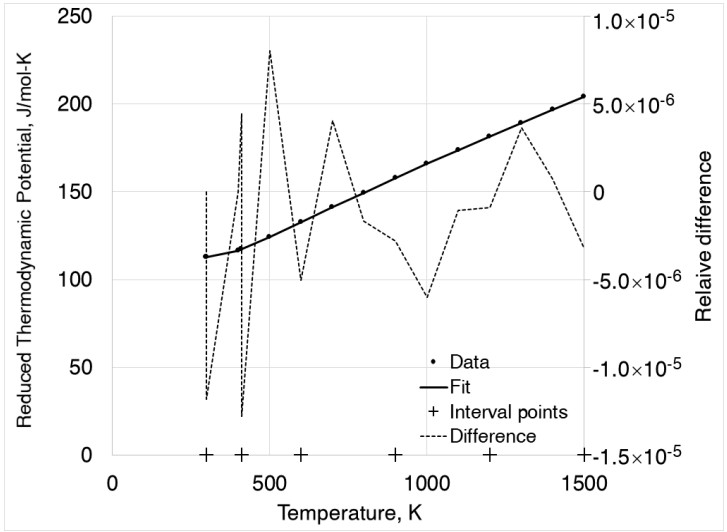
Ammonium Bromide (NH_4Br) – Br1H4N1(cr)

MW = 97.942 g/mol – $\Delta_f H^o$(298.15 K) = -271.542 kJ/mol – S^o(298.15 K) = 112.842 J/mol-K – p^o = 0.1 MPa



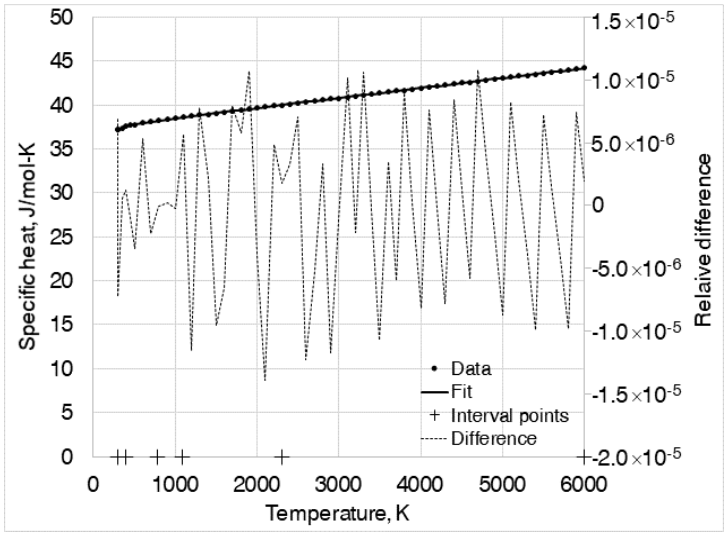


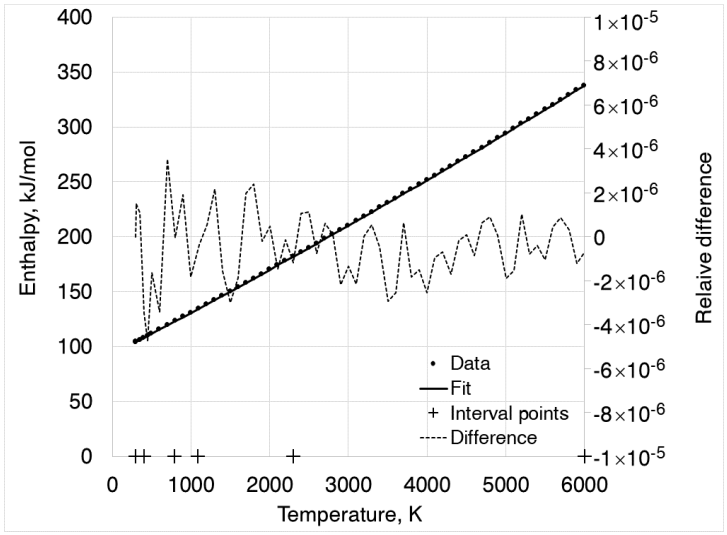


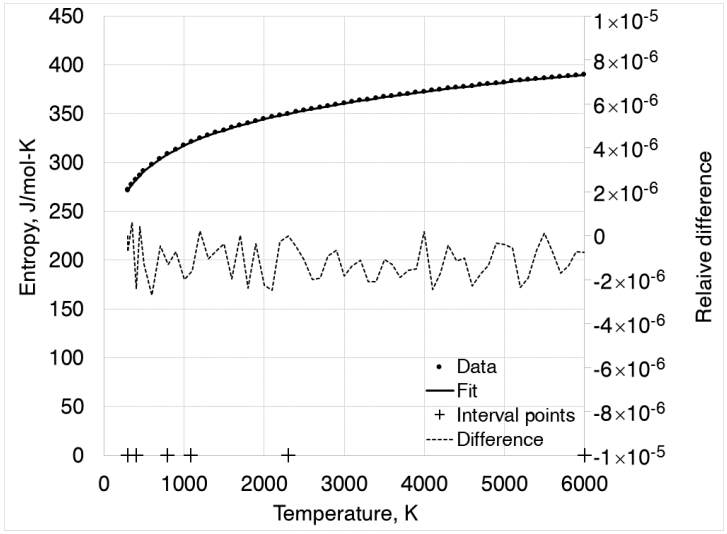


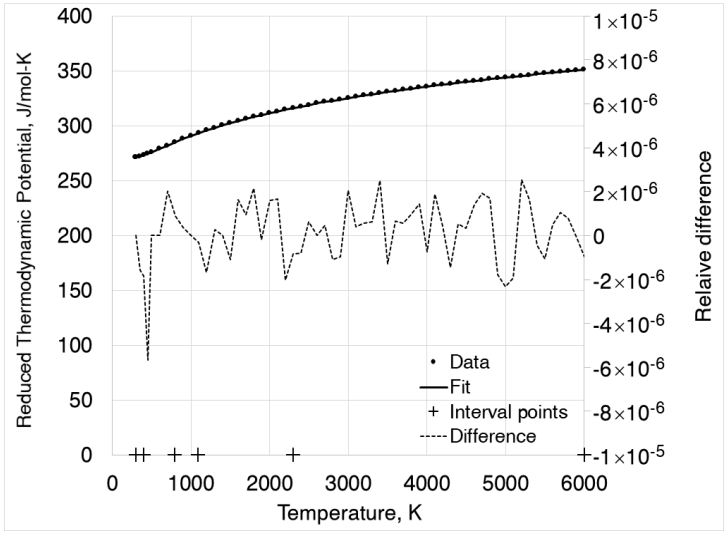
Mercury Bromide (HgBr) – Br1Hg1(g)

MW = 280.49 g/mol – $\Delta_f H^o$(298.15 K) = 104.182 kJ/mol – S^o(298.15 K) = 271.538 J/mol-K – p^o = 0.1 MPa



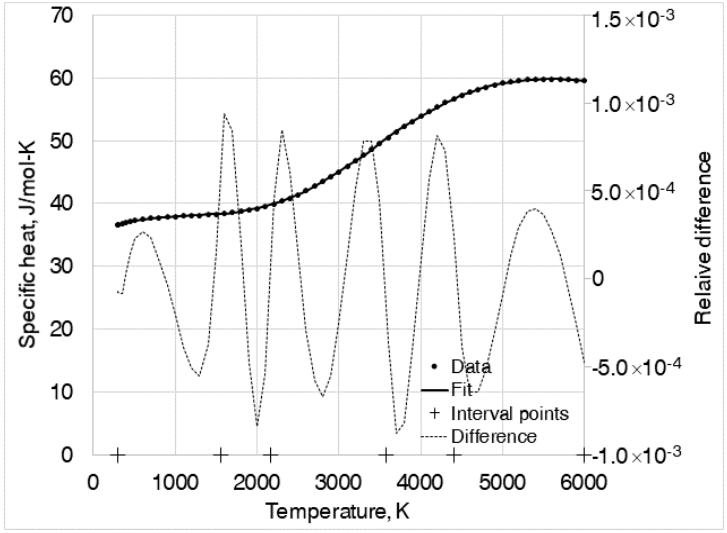


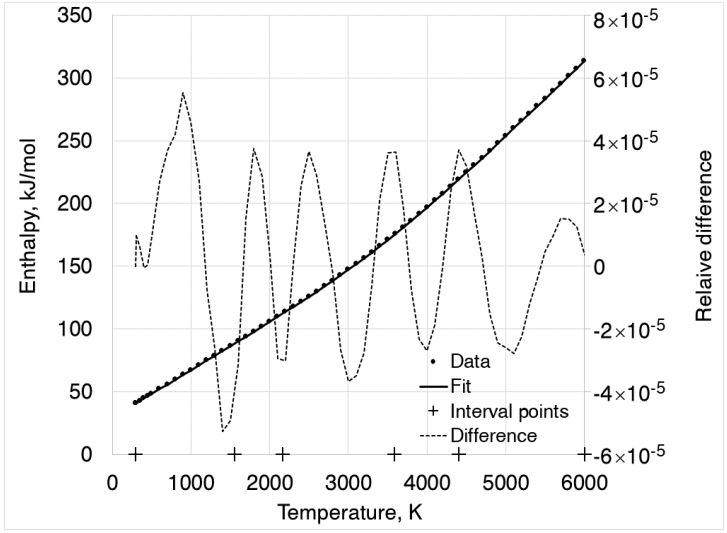


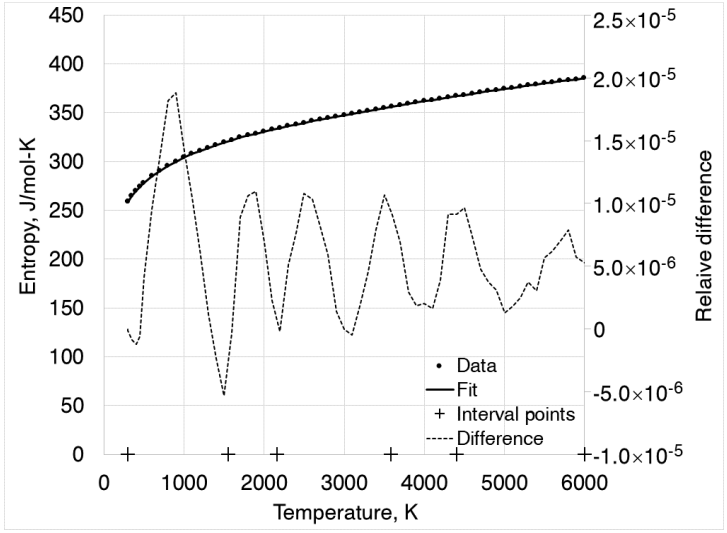


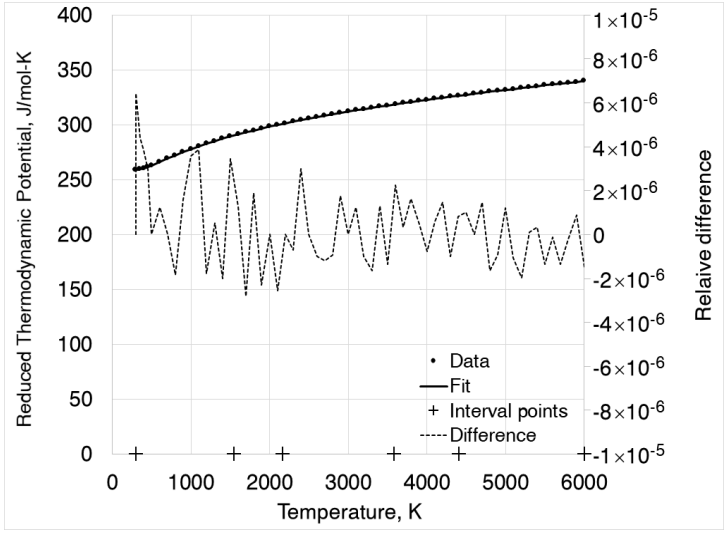
Iodine Bromide (IBr) – Br1I1(g)

MW = 206.808 g/mol – $\Delta_f H^o$(298.15 K) = 40.878 kJ/mol – S^o(298.15 K) = 258.952 J/mol-K – p^o = 0.1 MPa



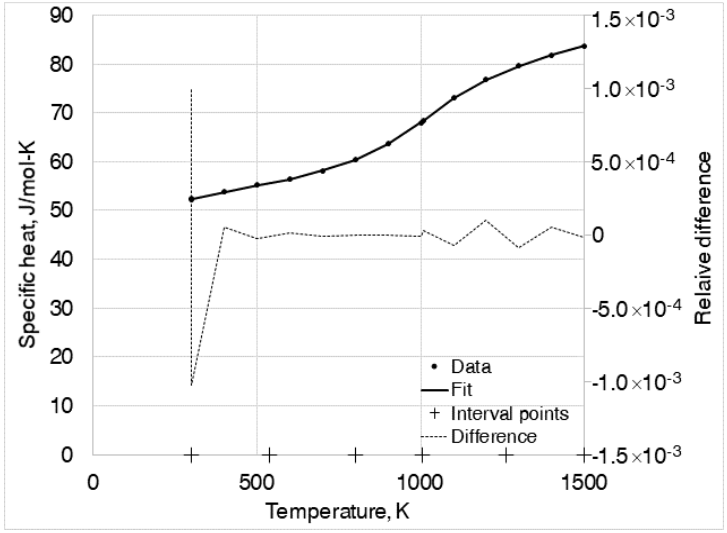


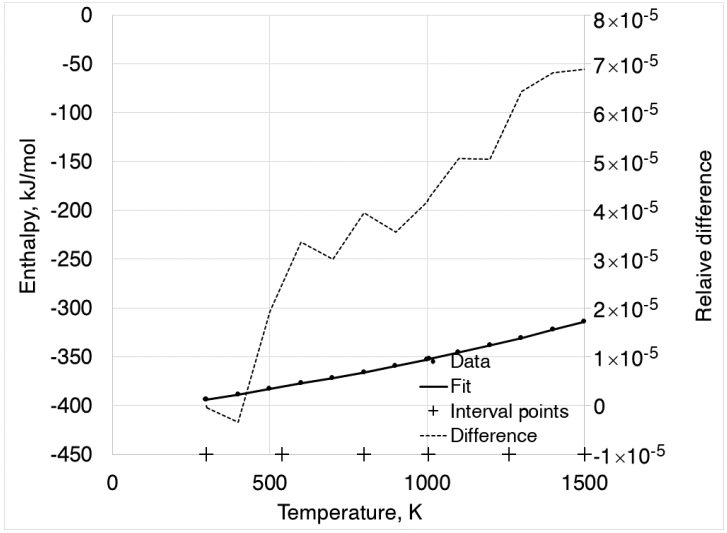


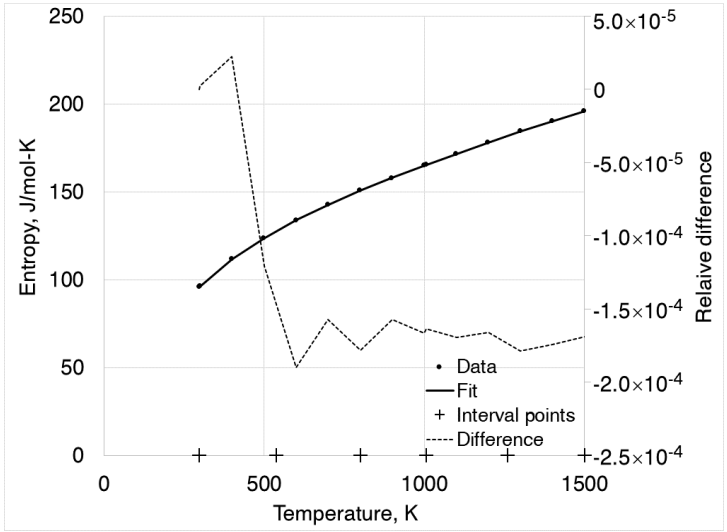


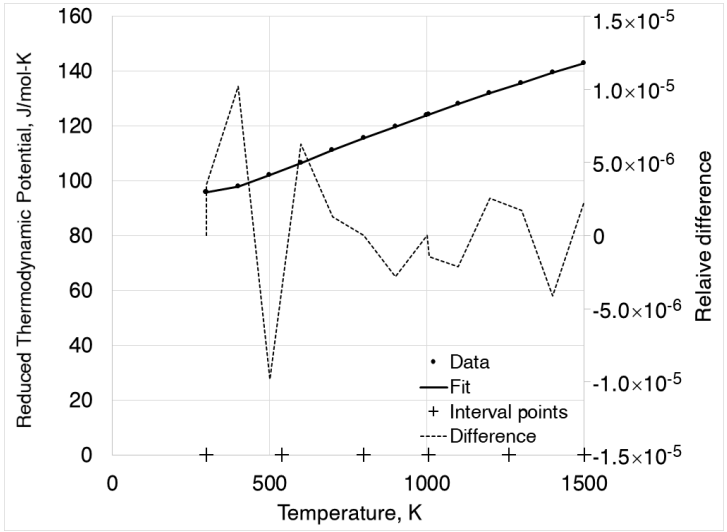
Potassium Bromide (KBr) – Br1K1(cr)

MW = 119.002 g/mol – $\Delta_f H^o$(298.15 K) = -393.798 kJ/mol – S^o(298.15 K) = 95.939 J/mol-K – p^o = 0.1 MPa



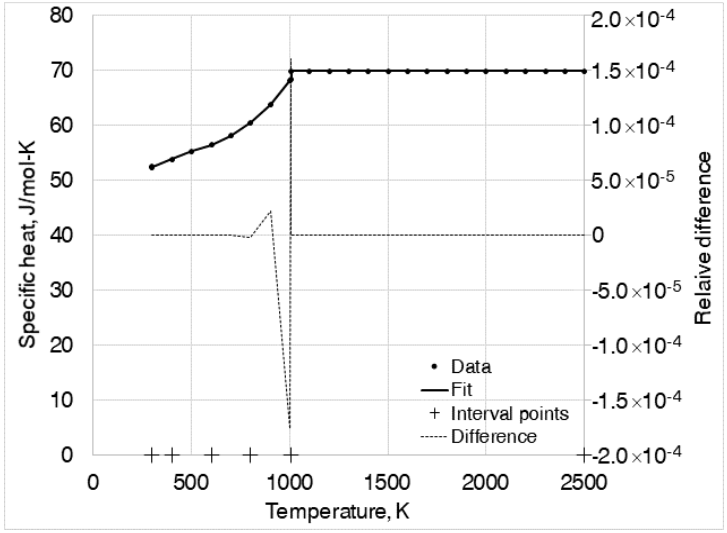


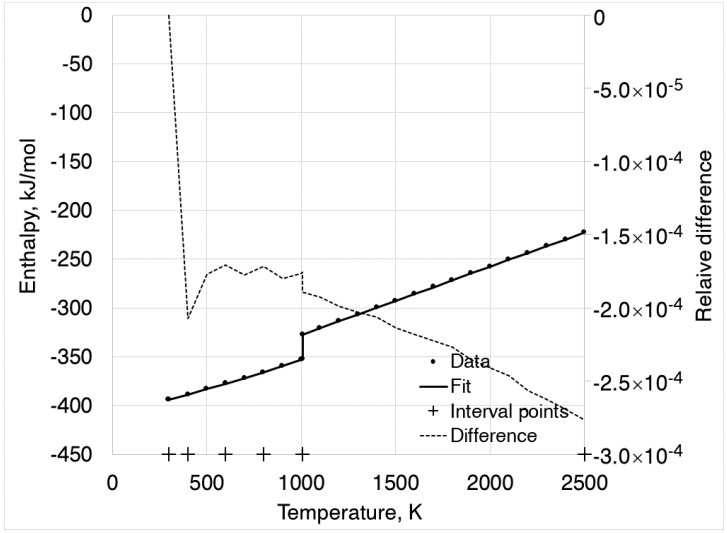


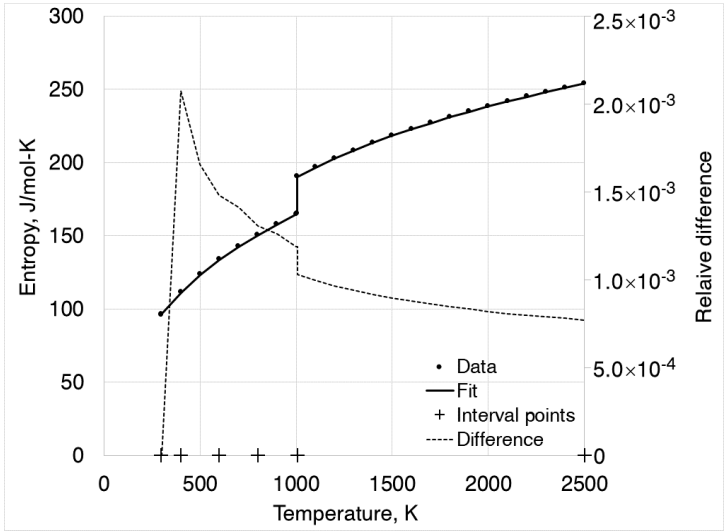


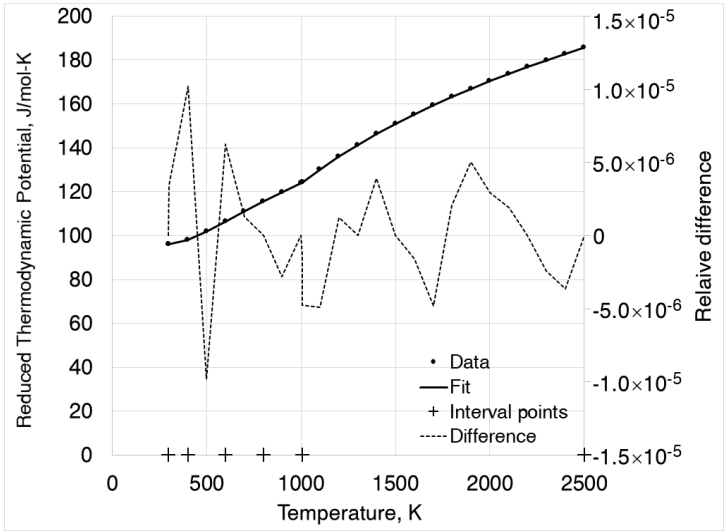
Potassium Bromide (KBr) – Br1K1(cr,l)

MW = 119.002 g/mol – $\Delta_f H^o$(298.15 K) = -393.798 kJ/mol – S^o(298.15 K) = 95.939 J/mol-K – p^o = 0.1 MPa



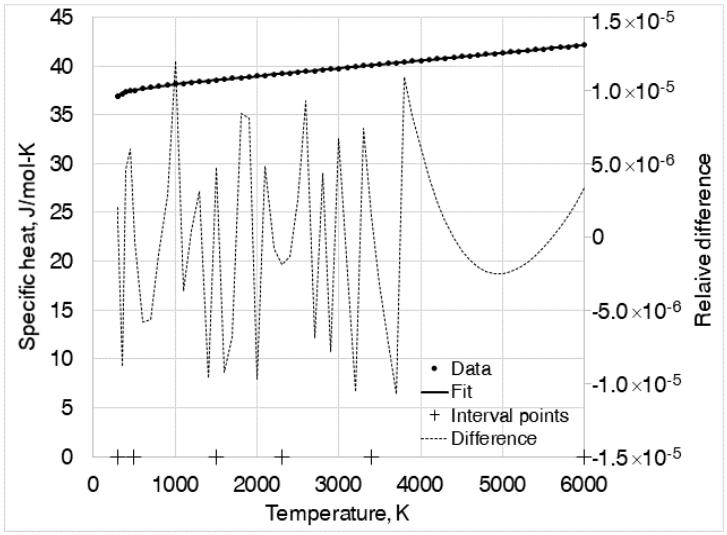


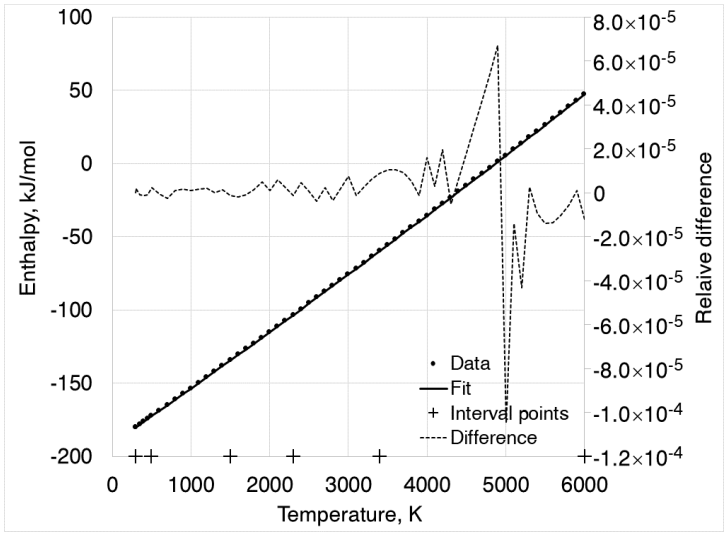


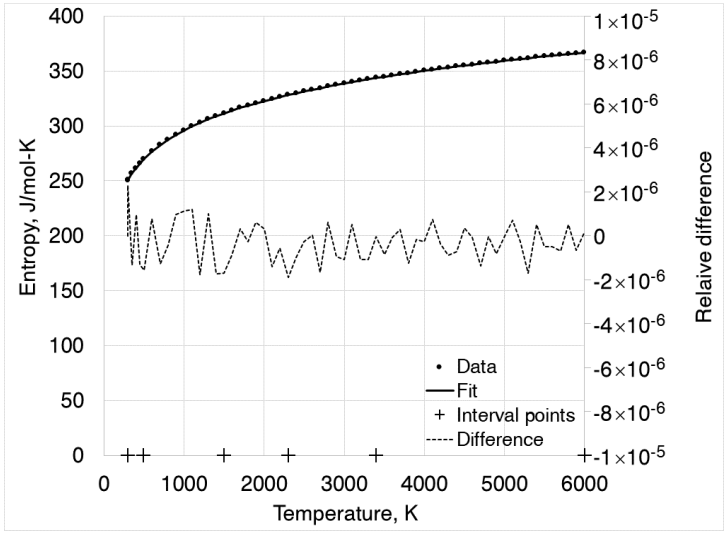


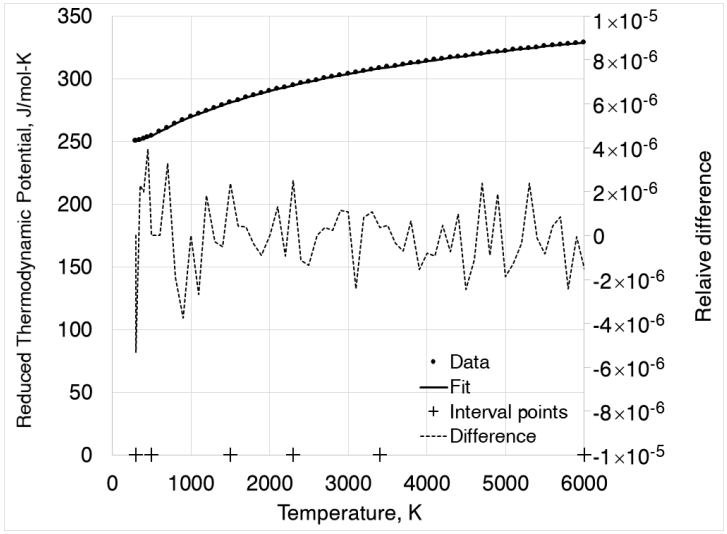
Potassium Bromide (KBr) – Br1K1(g)

MW = 119.002 g/mol – $\Delta_f H^o$(298.15 K) = -180.079 kJ/mol – S^o(298.15 K) = 250.532 J/mol-K – p^o = 0.1 MPa



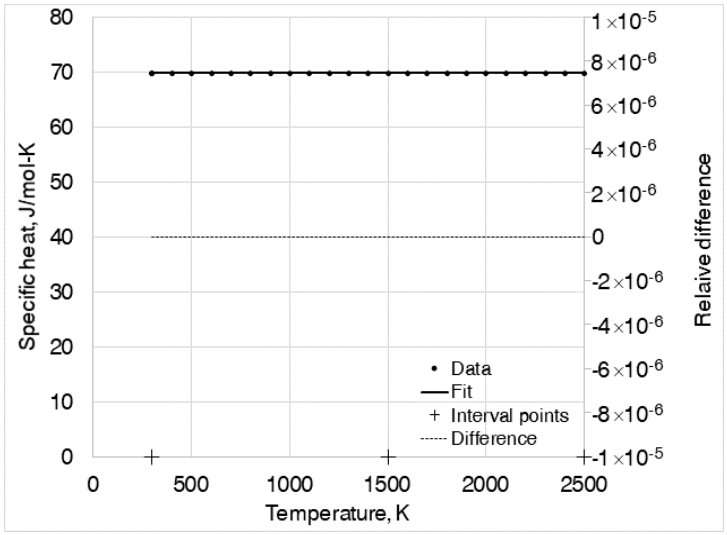


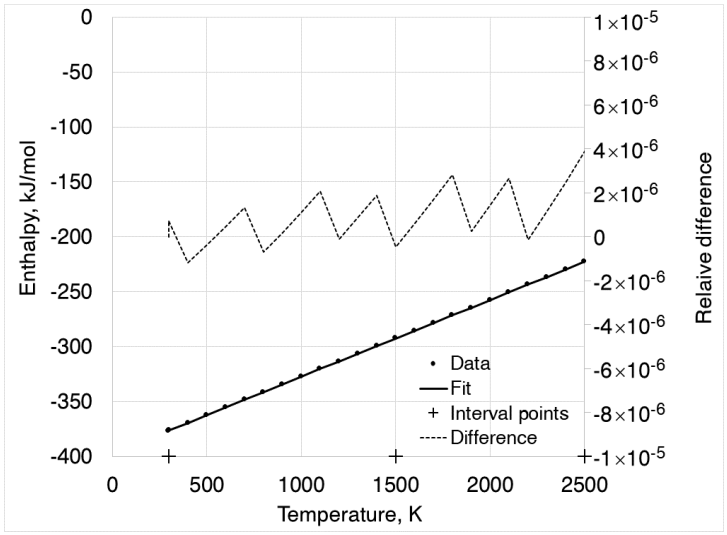


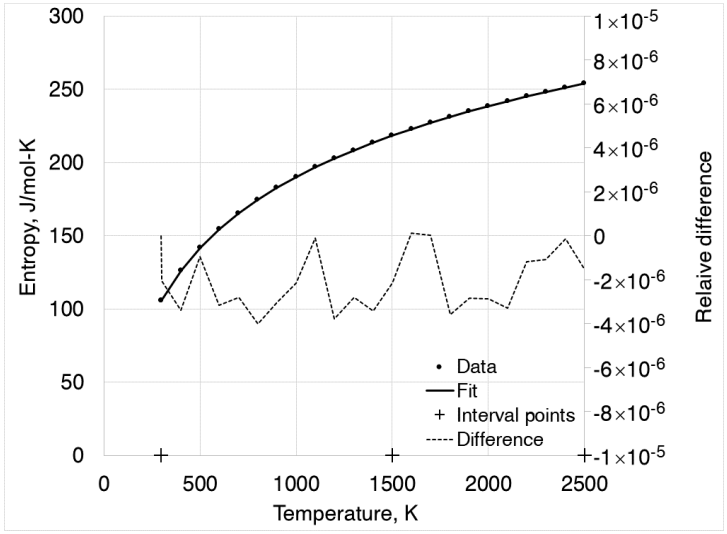


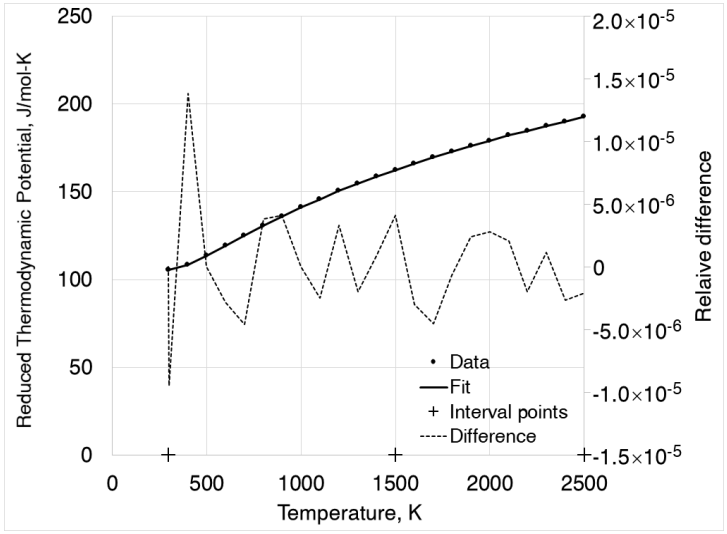
Potassium Bromide (KBr) – Br1K1(l)

MW = 119.002 g/mol – $\Delta_f H^o$(298.15 K) = -376.492 kJ/mol – S^o(298.15 K) = 105.54 J/mol-K – p^o = 0.1 MPa



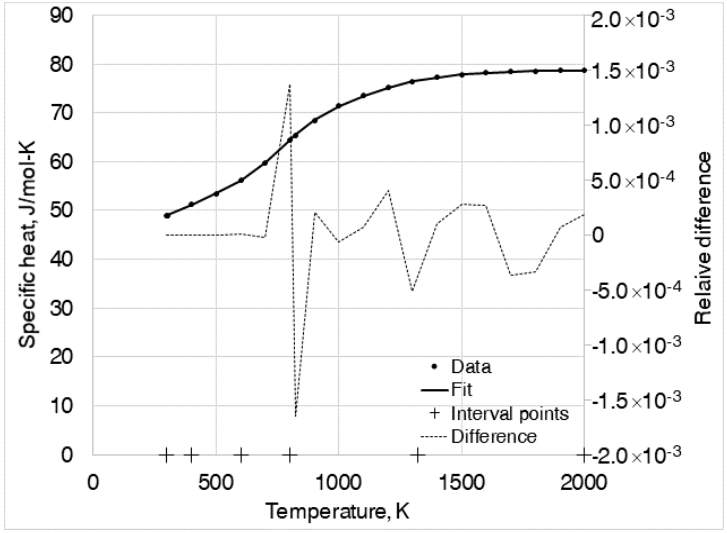


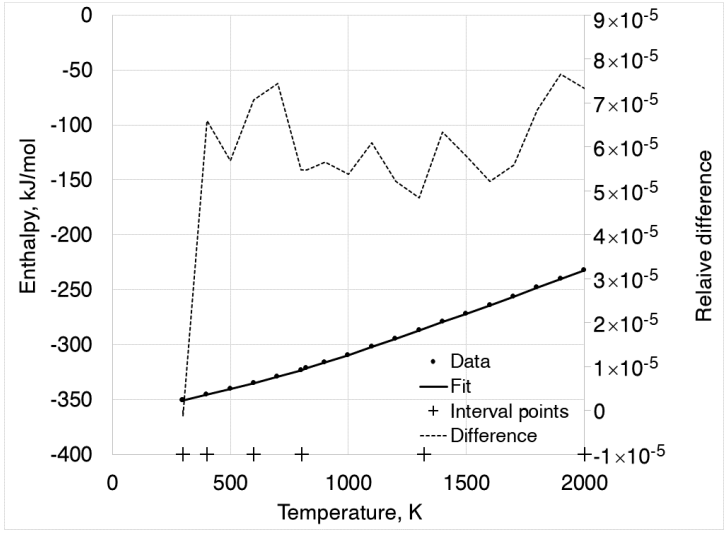


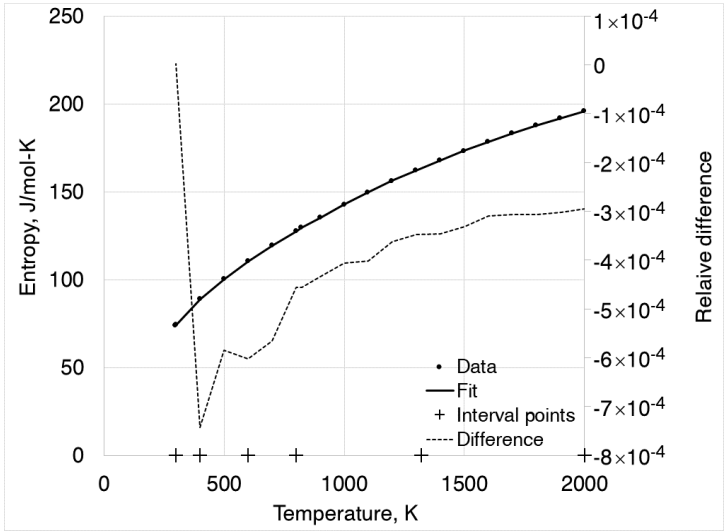


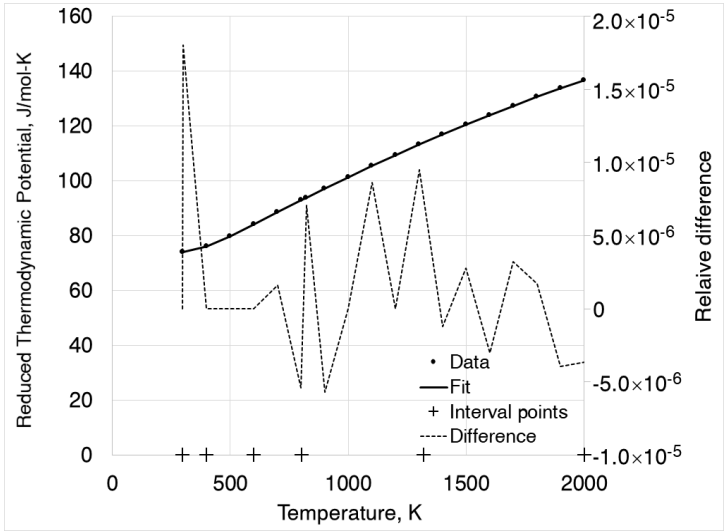
Lithium Bromide (LiBr) – Br1Li1(cr)

MW = 86.845 g/mol – $\Delta_f H^o$(298.15 K) = -350.912 kJ/mol – S^o(298.15 K) = 74.057 J/mol-K – p^o = 0.1 MPa



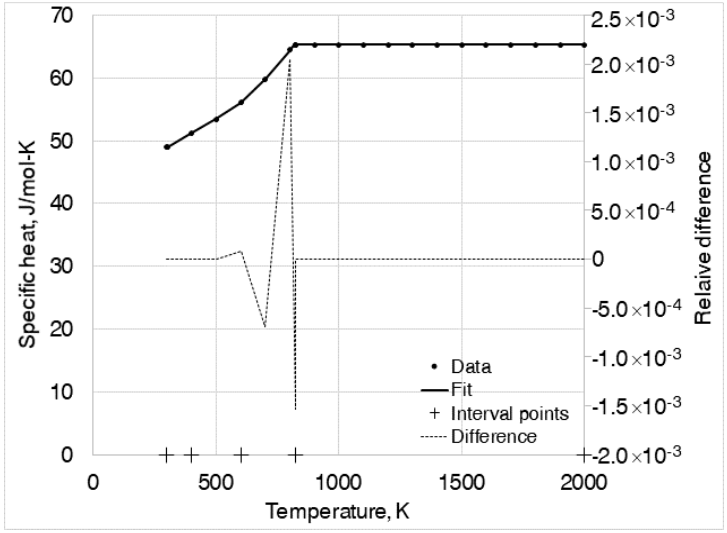


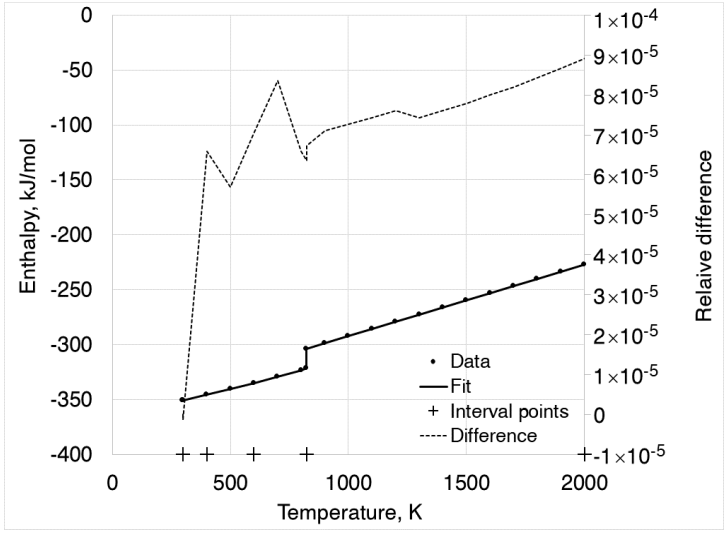


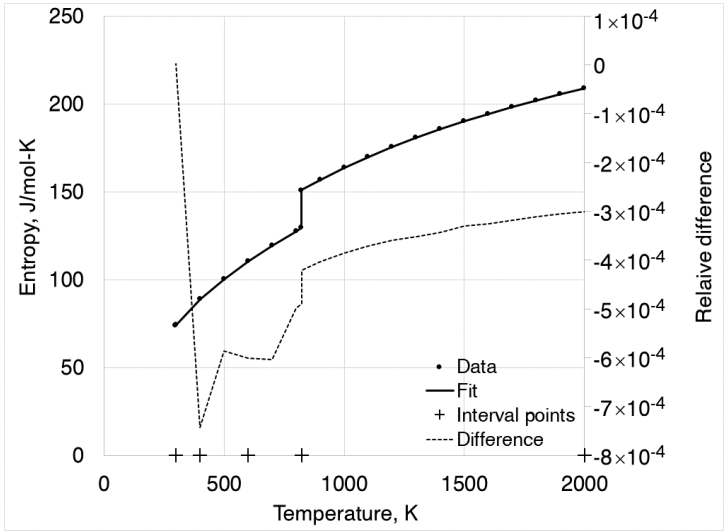


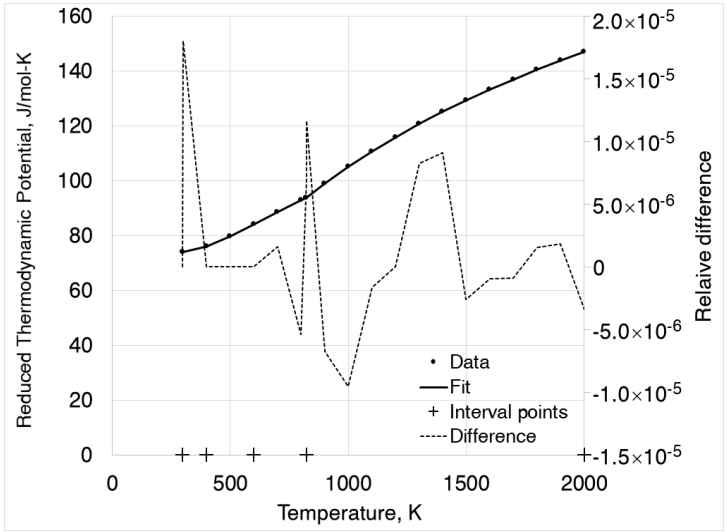
Lithium Bromide (LiBr) – Br1Li1(cr,l)

MW = 86.845 g/mol – $\Delta_f H^o$(298.15 K) = -350.912 kJ/mol – S^o(298.15 K) = 74.057 J/mol-K – p^o = 0.1 MPa



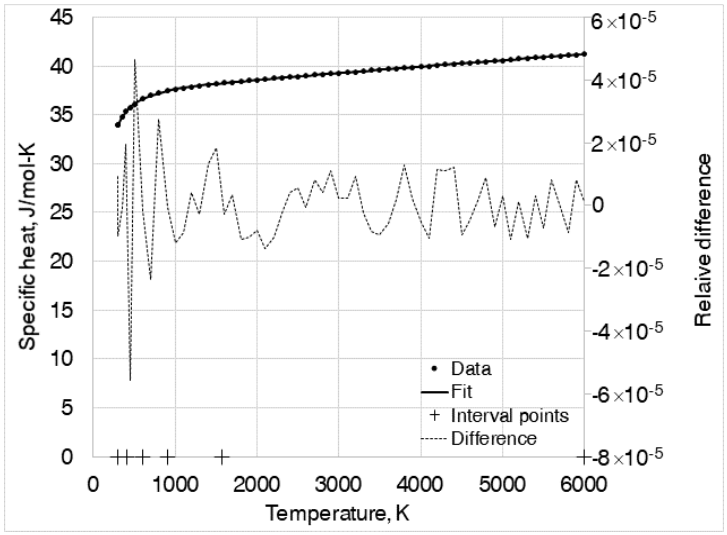


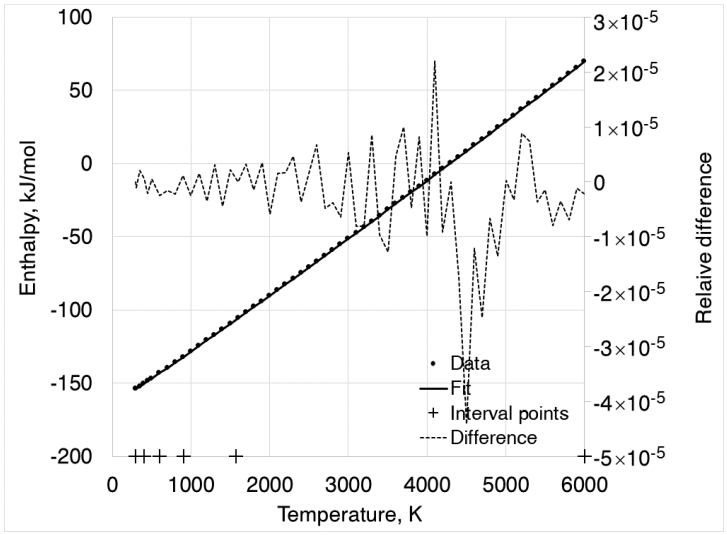


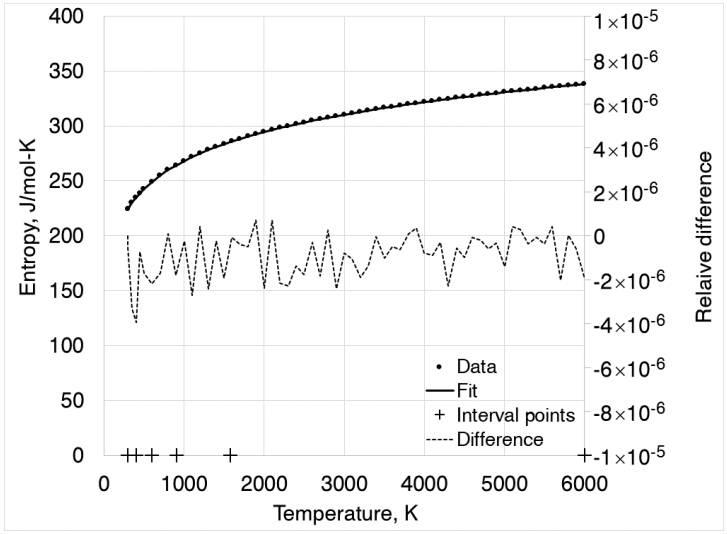


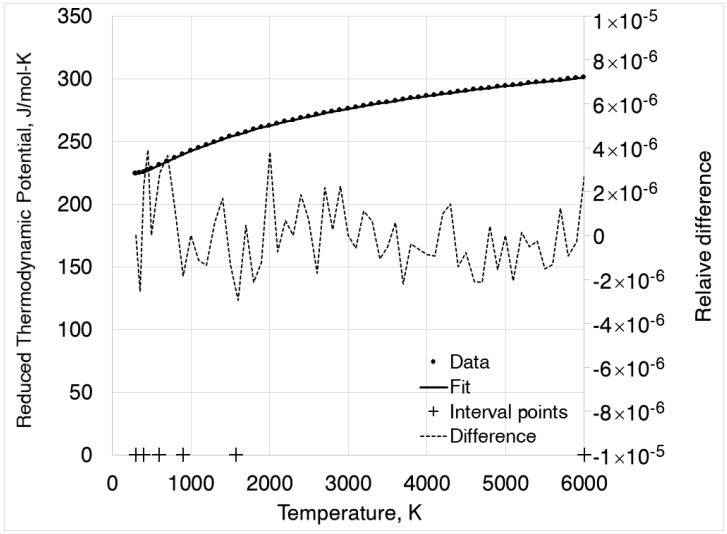
Lithium Bromide (LiBr) – Br1Li1(g)

MW = 86.845 g/mol – $\Delta_f H^o$(298.15 K) = -153.971 kJ/mol – S^o(298.15 K) = 224.345 J/mol-K – p^o = 0.1 MPa



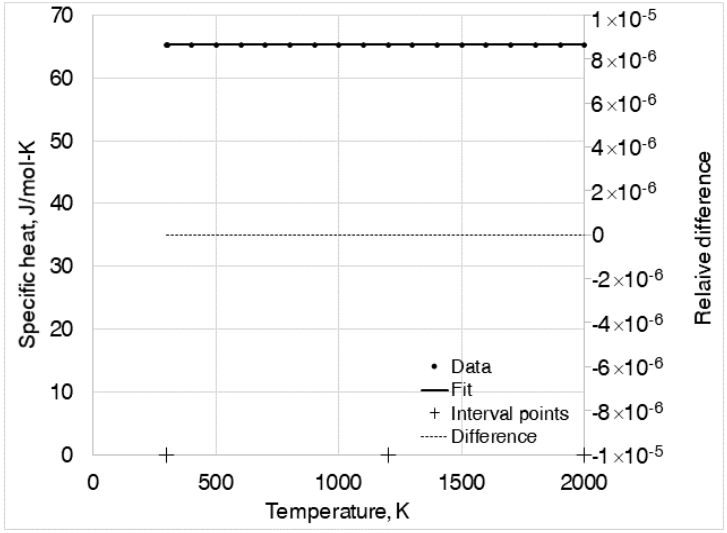


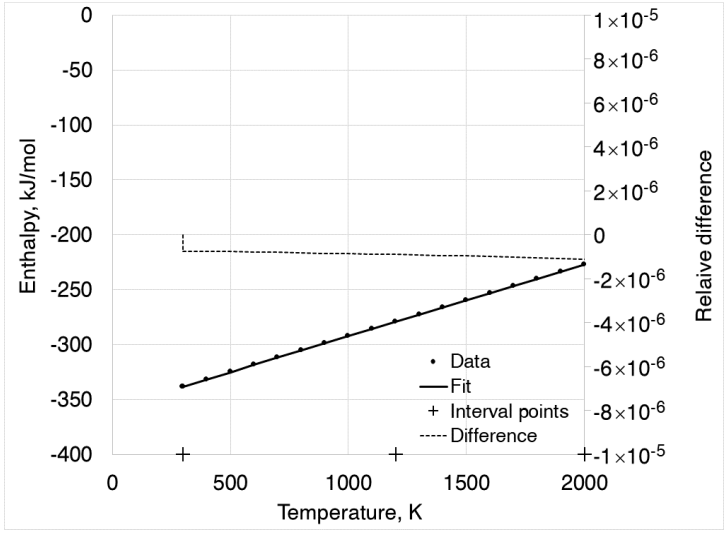


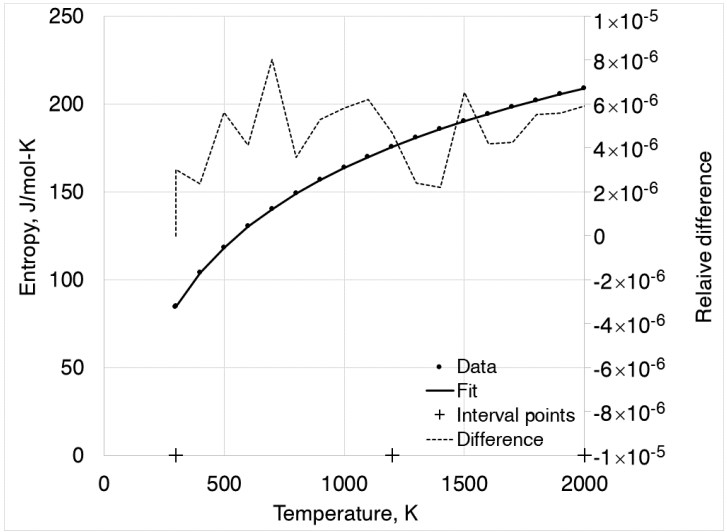


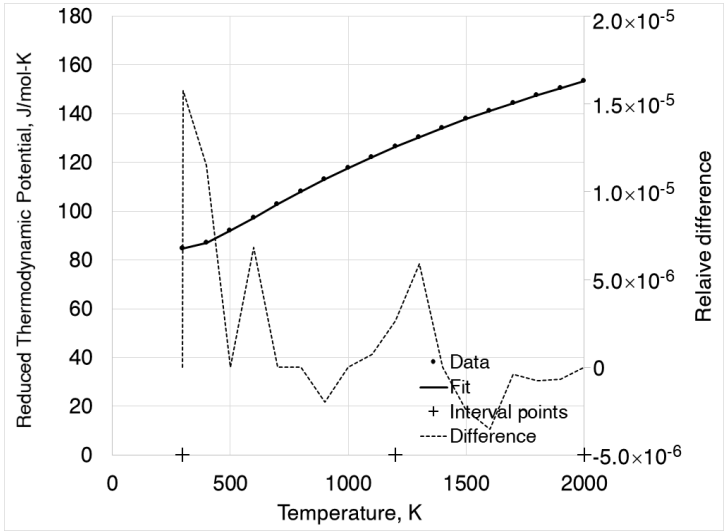
Lithium Bromide (LiBr) – Br1Li1(l)

MW = 86.845 g/mol – $\Delta_f H^o$(298.15 K) = -338.226 kJ/mol – S^o(298.15 K) = 84.603 J/mol-K – p^o = 0.1 MPa



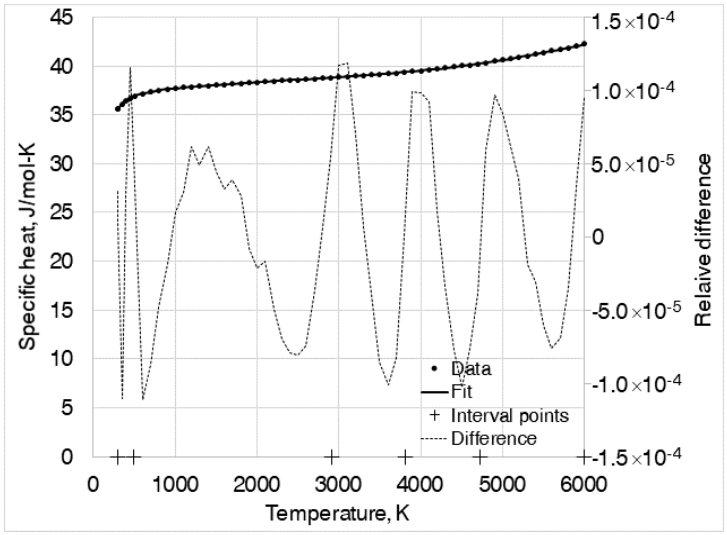


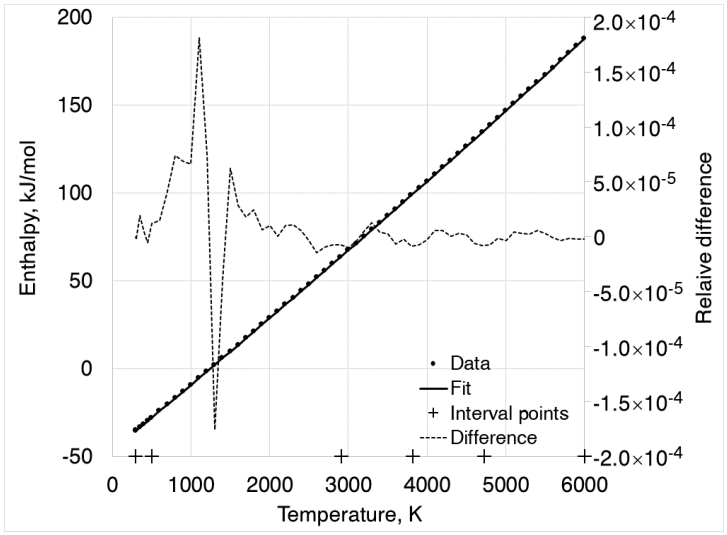


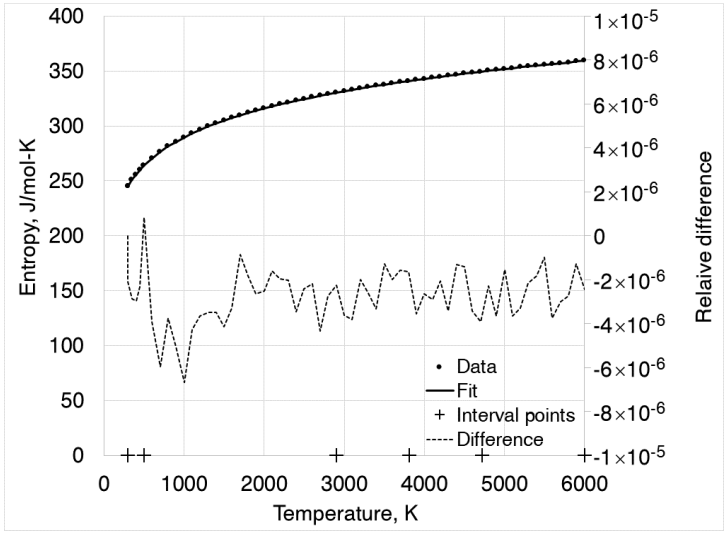


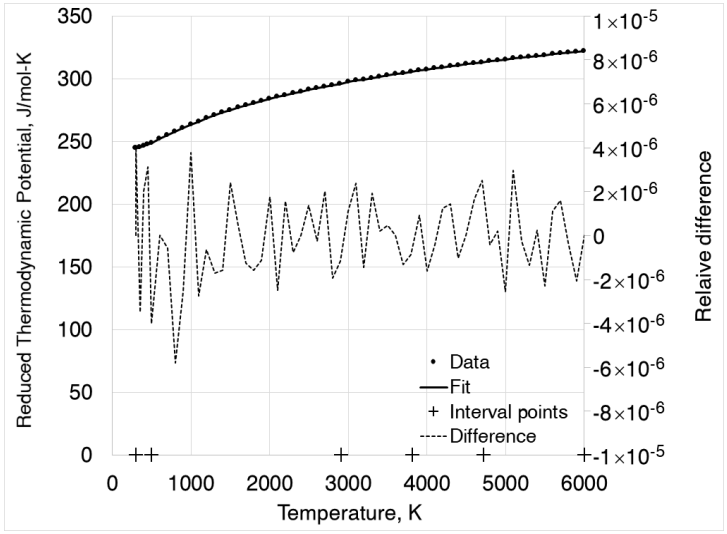
Magnesium Bromide (MgBr) – Br1Mg1(g)

MW = 104.209 g/mol – $\Delta_f H^o$(298.15 K) = -35.342 kJ/mol – S^o(298.15 K) = 244.946 J/mol-K – p^o = 0.1 MPa



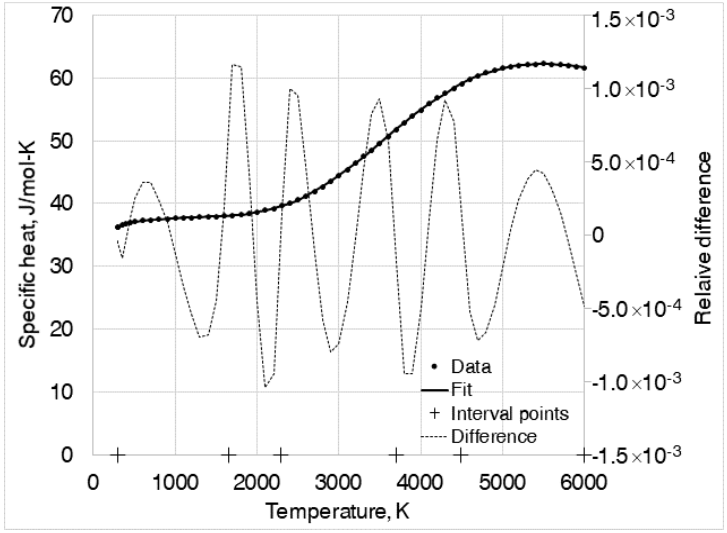


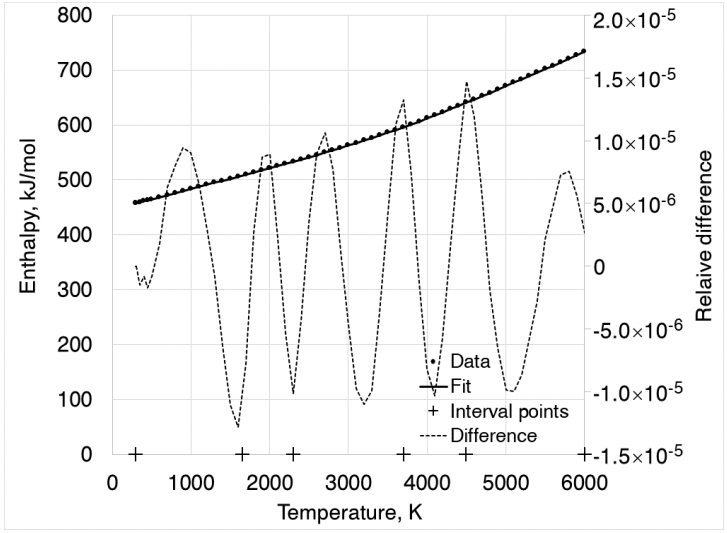


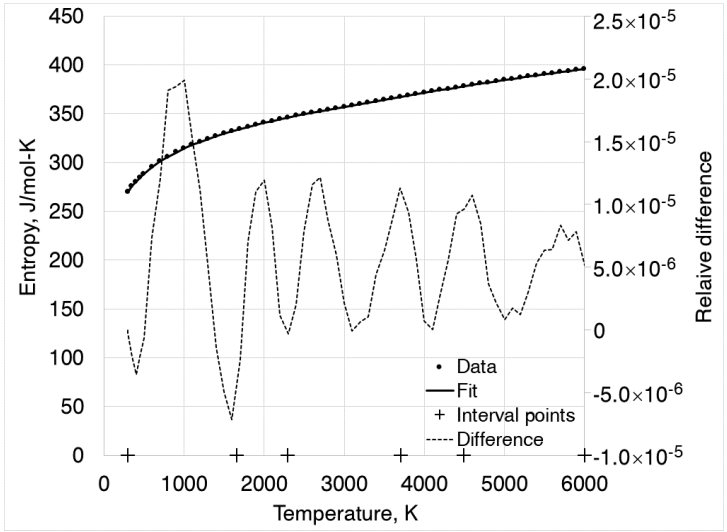


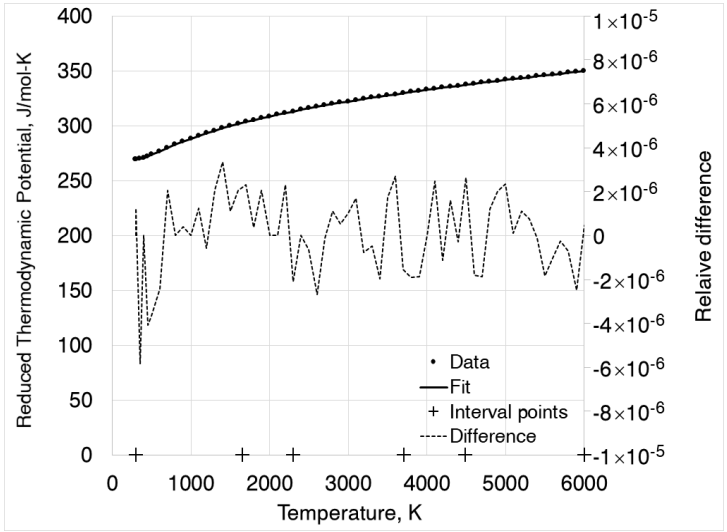
Molybdenum Bromide (MoBr) – Br1Mo1(g)

MW = 175.84 g/mol – $\Delta_f H^o$(298.15 K) = 457.311 kJ/mol – S^o(298.15 K) = 269.523 J/mol-K – p^o = 0.1 MPa



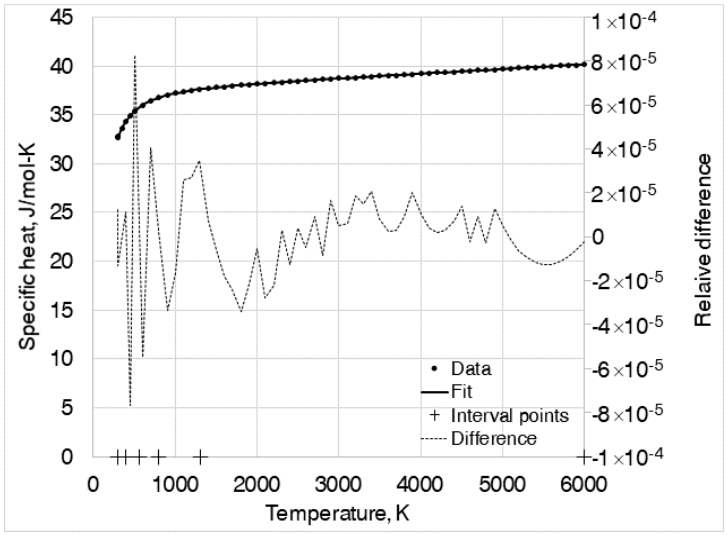


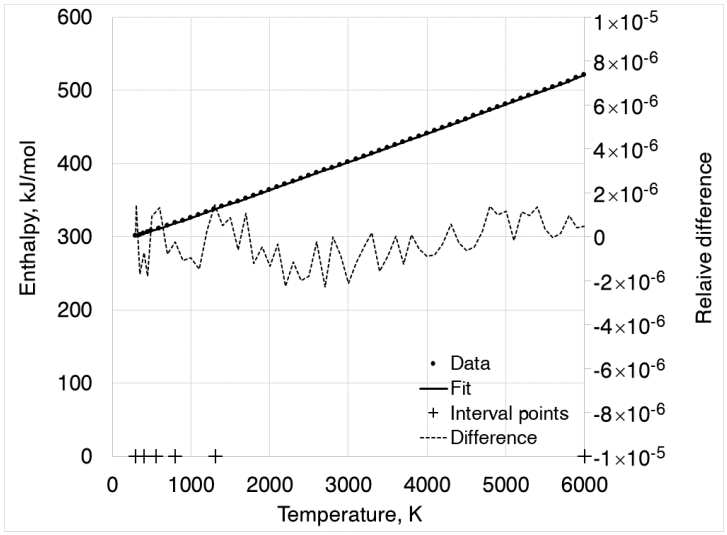


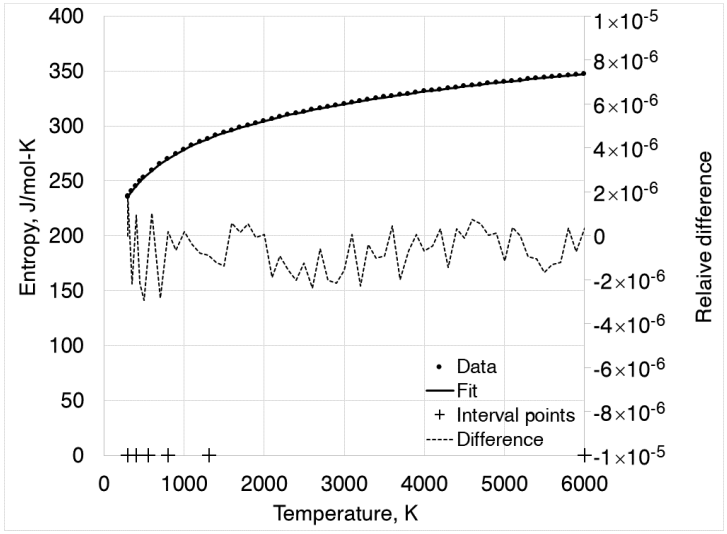


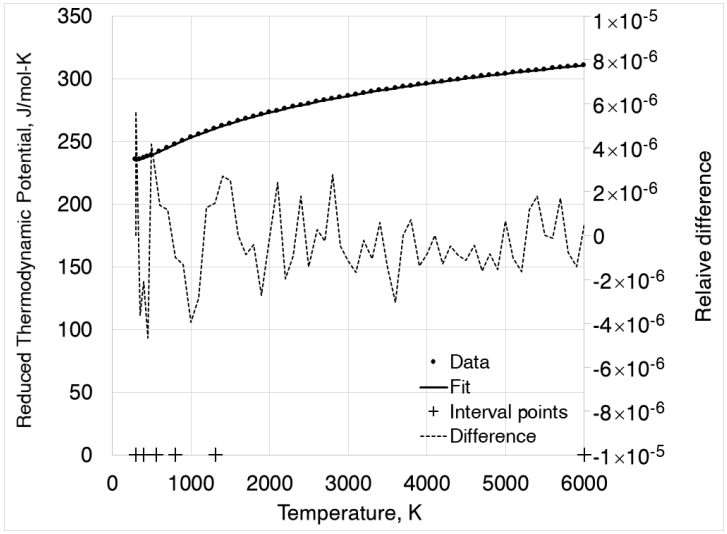
Bromoimidogen (NBr) – Br1N1(g)

MW = 93.911 g/mol – $\Delta_f H^o$(298.15 K) = 300.758 kJ/mol – S^o(298.15 K) = 235.482 J/mol-K – p^o = 0.1 MPa



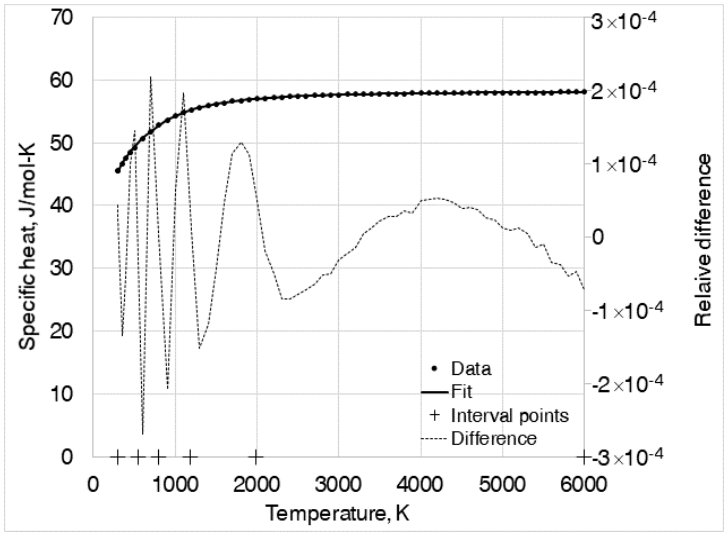


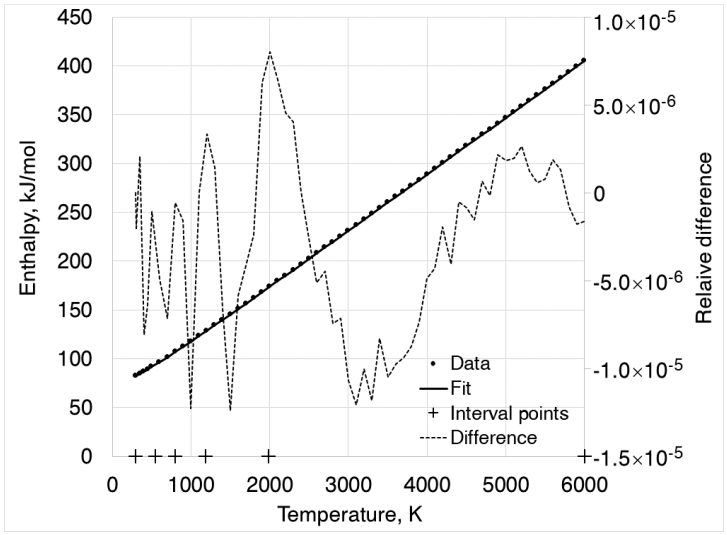


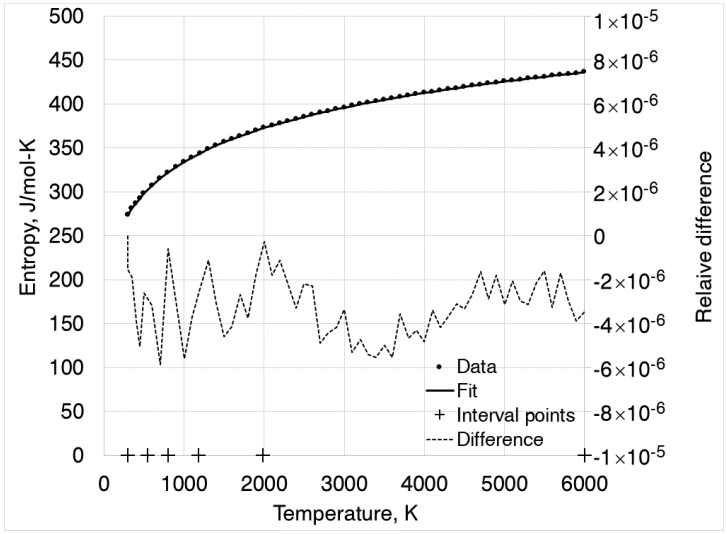


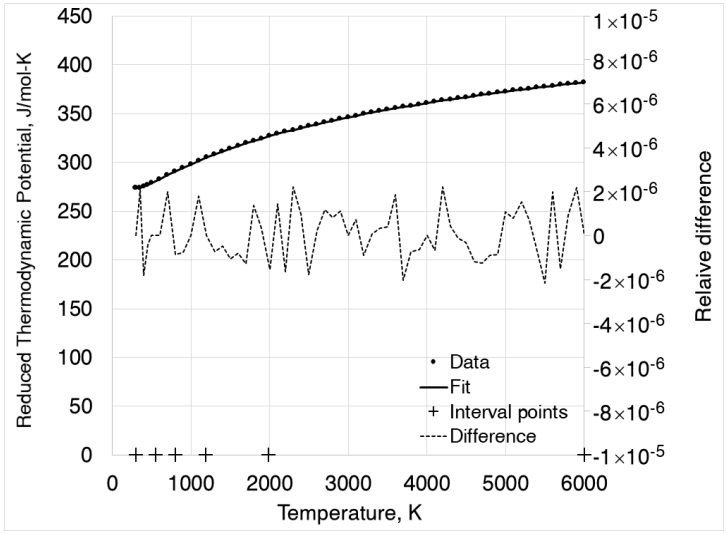
Nitrosyl Bromide (ONBr) – Br1N1O1(g)

MW = 109.91 g/mol – $\Delta_f H^o$(298.15 K) = 82.132 kJ/mol – S^o(298.15 K) = 273.525 J/mol-K – p^o = 0.1 MPa



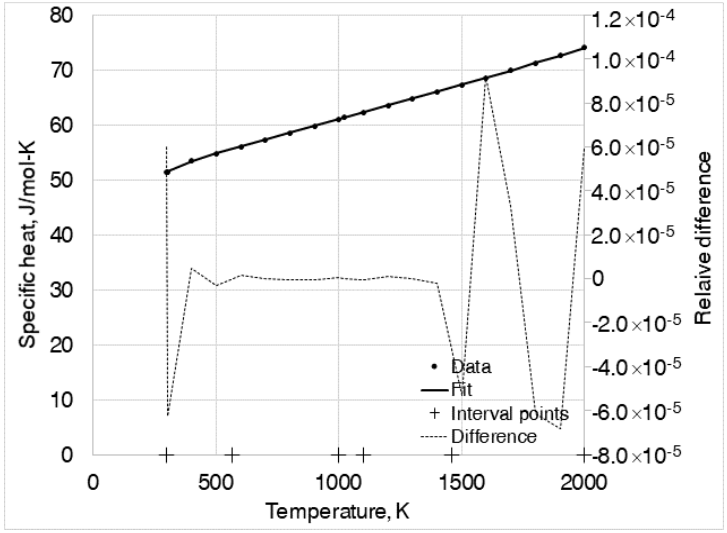


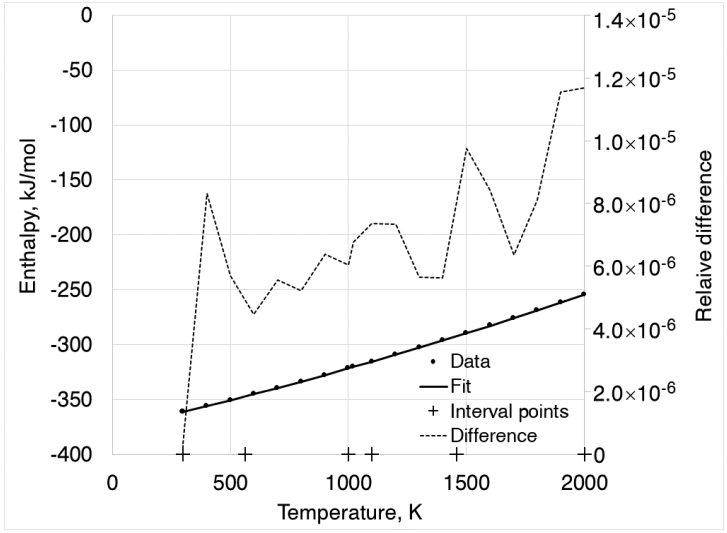


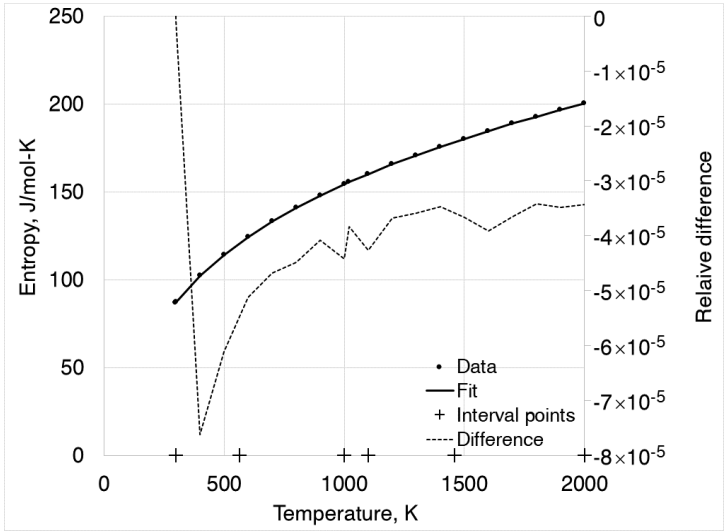


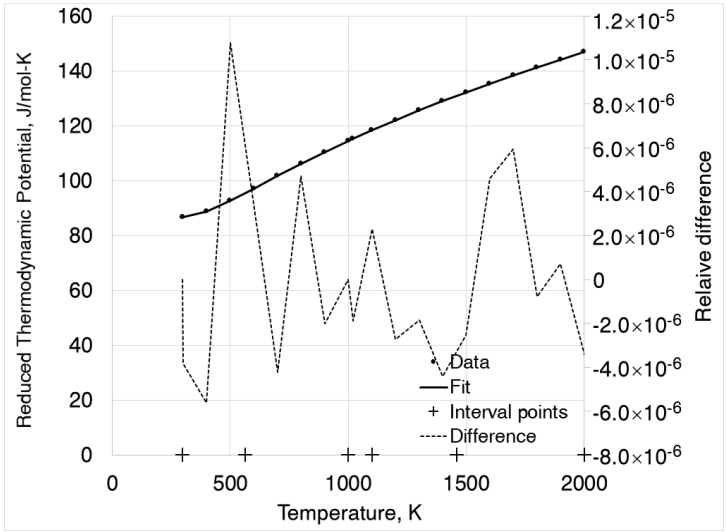
Sodium Bromide (NaBr) – Br1Na1(cr)

MW = 102.894 g/mol – $\Delta_f H^o$(298.15 K) = -361.414 kJ/mol – S^o(298.15 K) = 86.818 J/mol-K – p^o = 0.1 MPa



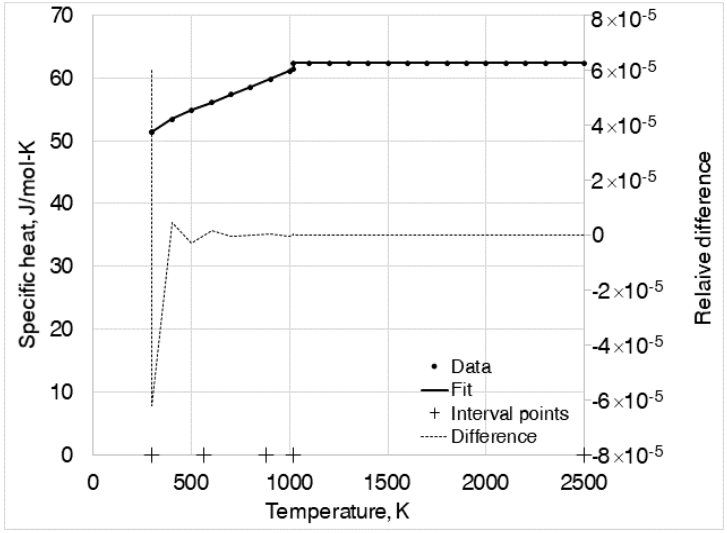


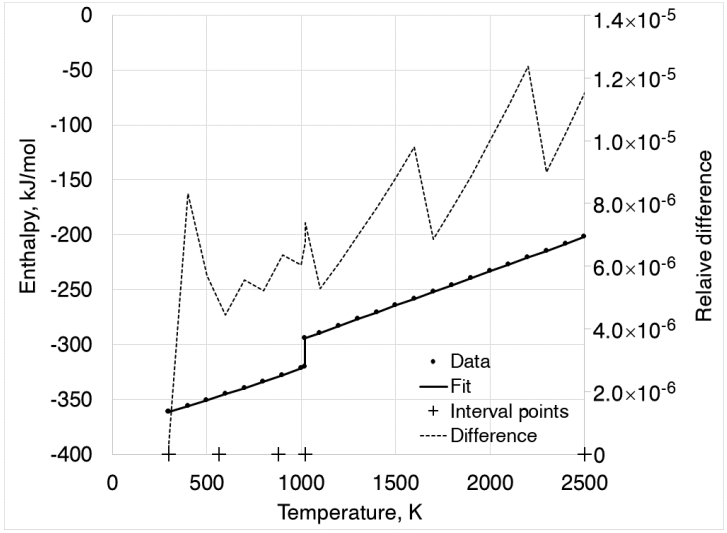


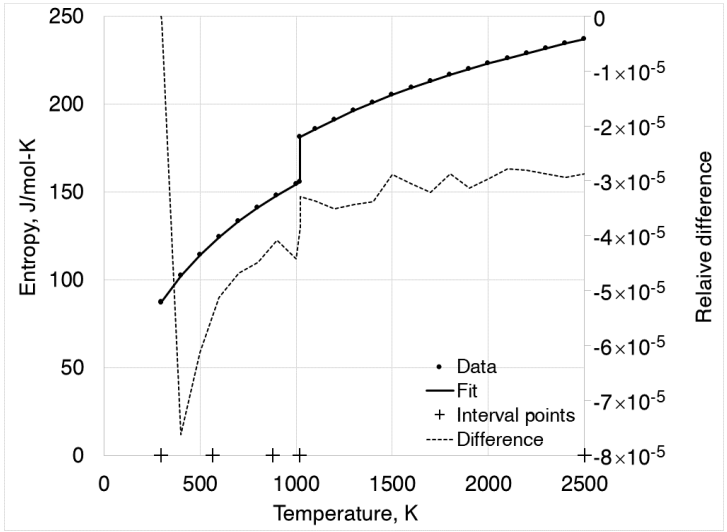


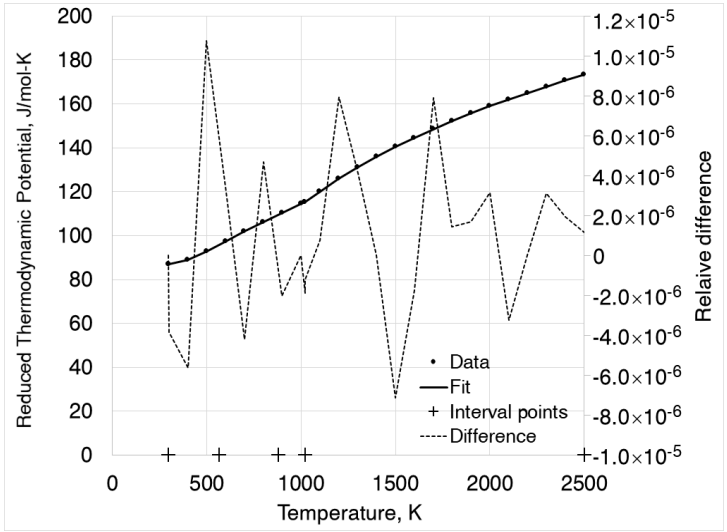
Sodium Bromide (NaBr) – Br1Na1(cr,l)

MW = 102.894 g/mol – $\Delta_f H^o$(298.15 K) = -361.414 kJ/mol – S^o(298.15 K) = 86.818 J/mol-K – p^o = 0.1 MPa



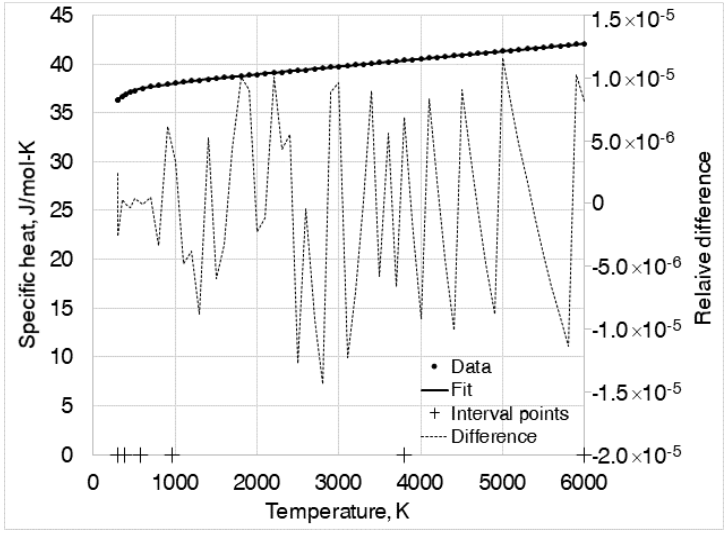


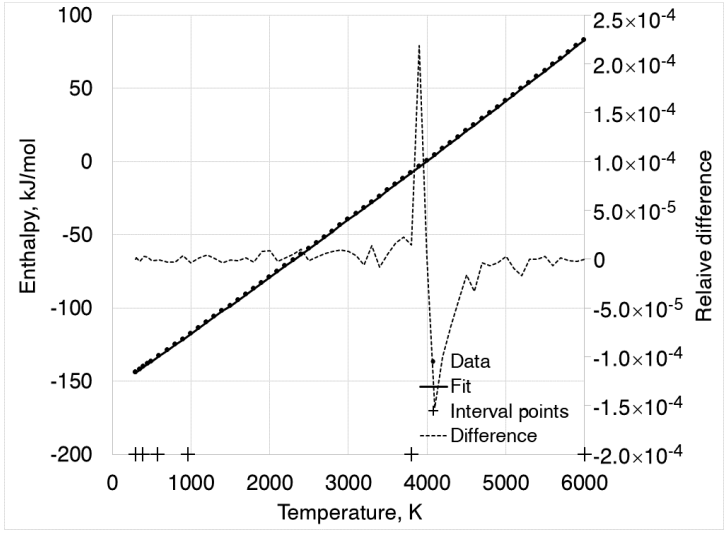


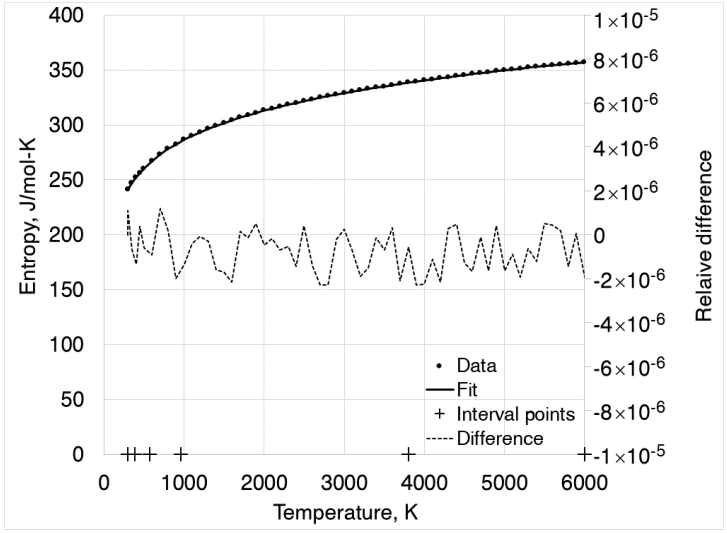


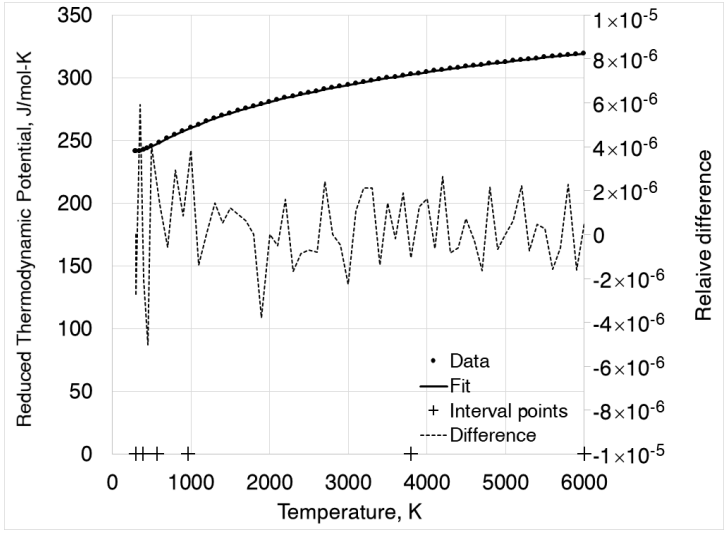
Sodium Bromide (NaBr) – Br1Na1(g)

MW = 102.894 g/mol – $\Delta_f H^o$(298.15 K) = -143.93 kJ/mol – S^o(298.15 K) = 241.218 J/mol-K – p^o = 0.1 MPa



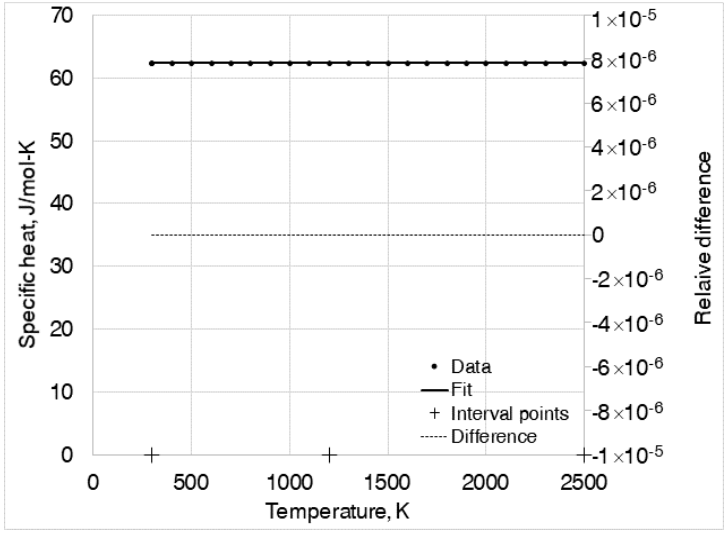


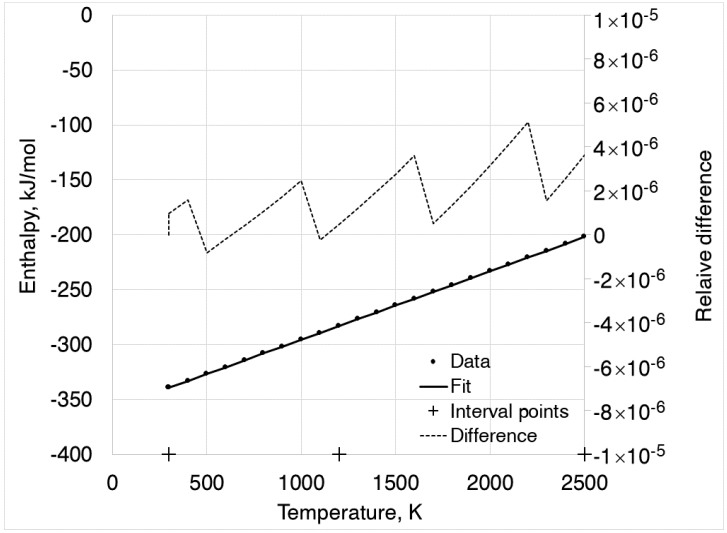


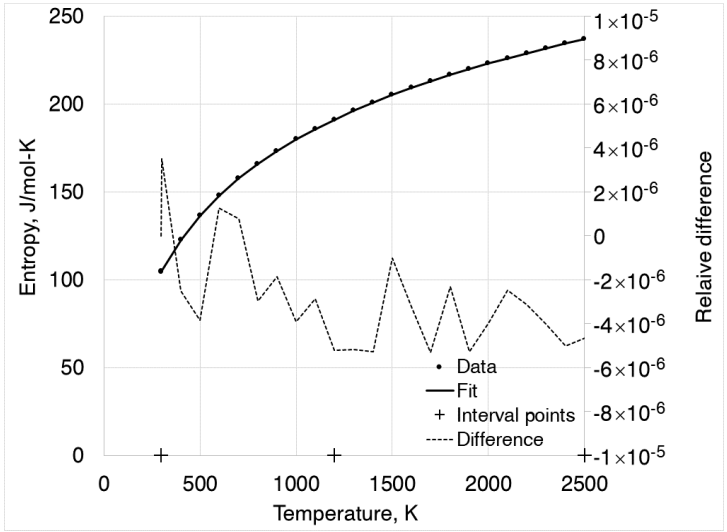


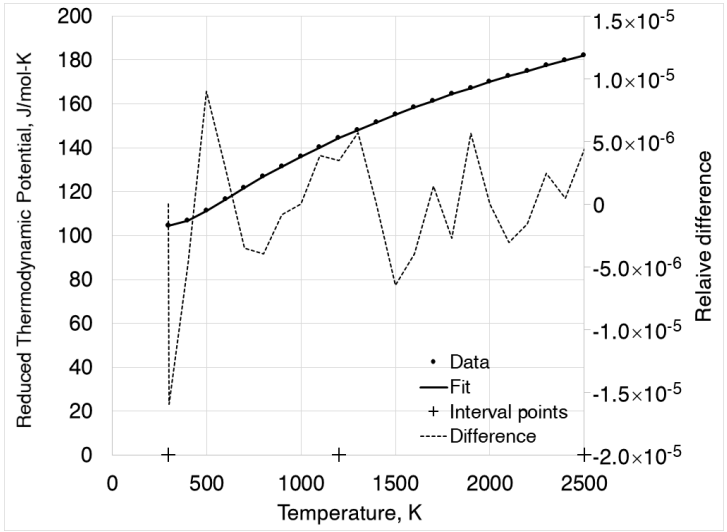
Sodium Bromide (NaBr) – Br1Na1(l)

MW = 102.894 g/mol – $\Delta_f H^o$(298.15 K) = -339.343 kJ/mol – S^o(298.15 K) = 104.363 J/mol-K – p^o = 0.1 MPa



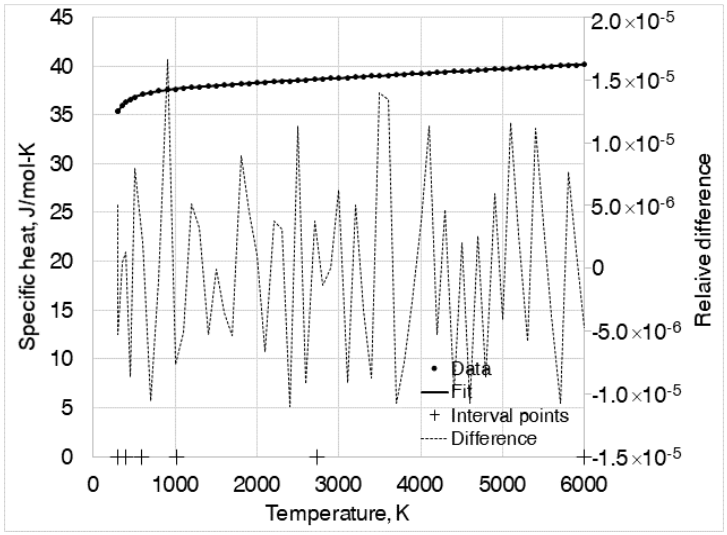


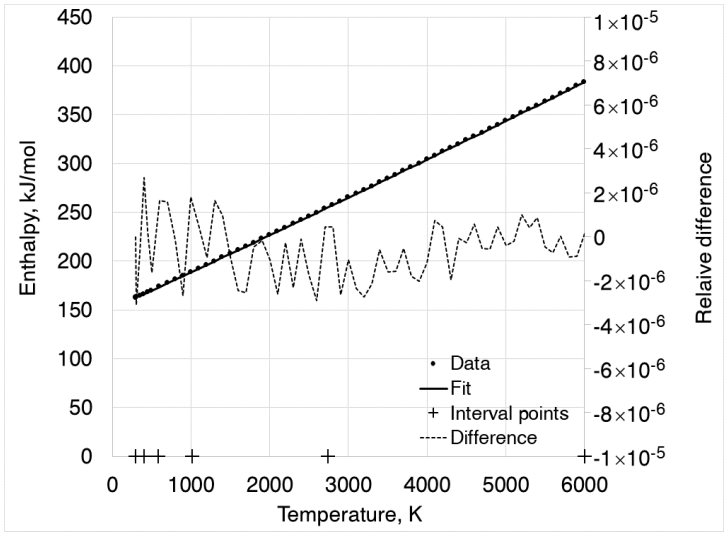


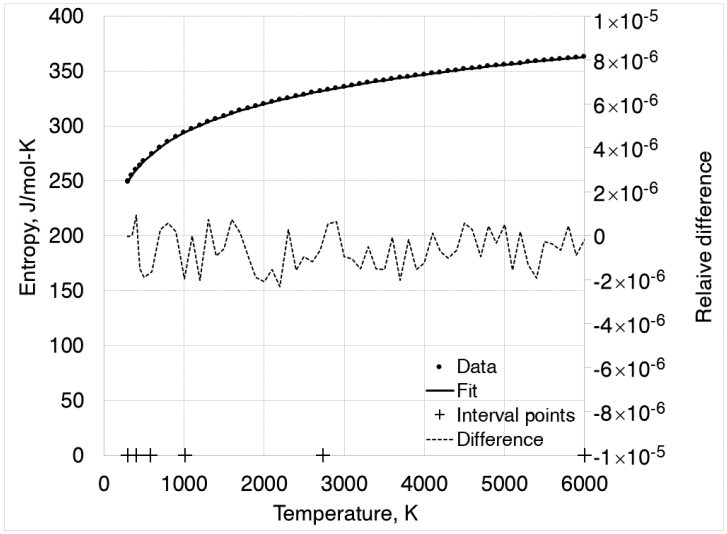


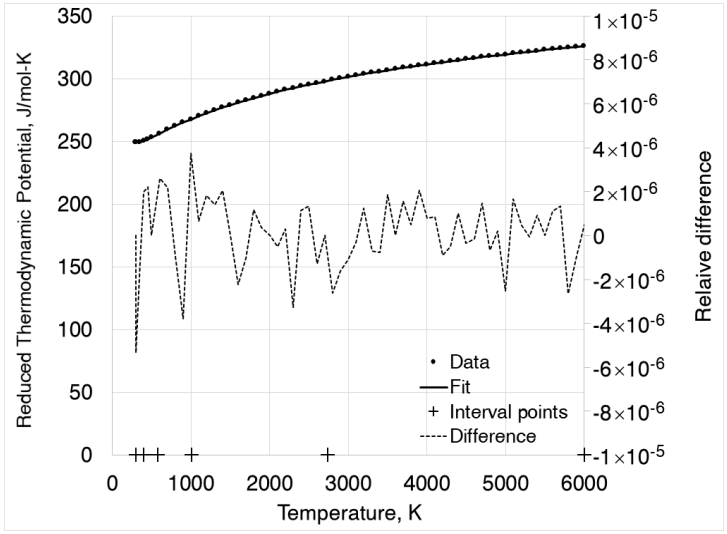
Phosphorus Bromide (PBr) – Br1P1(g)

MW = 110.878 g/mol – $\Delta_f H^o$(298.15 K) = 162.601 kJ/mol – S^o(298.15 K) = 249.204 J/mol-K – p^o = 0.1 MPa



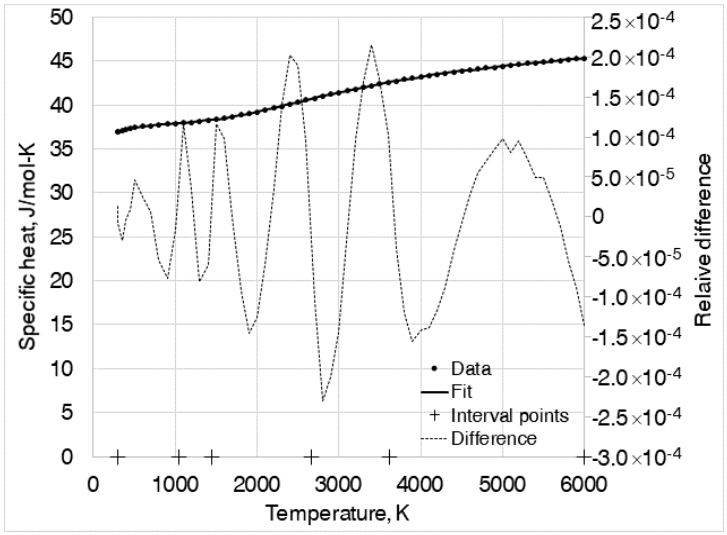


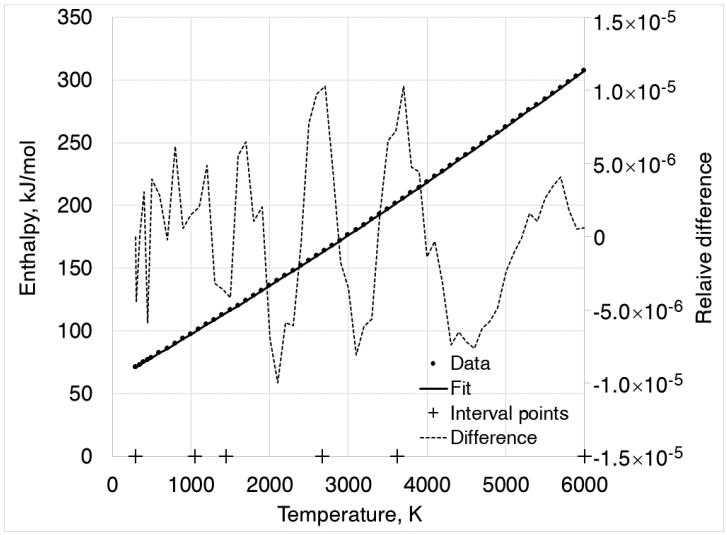


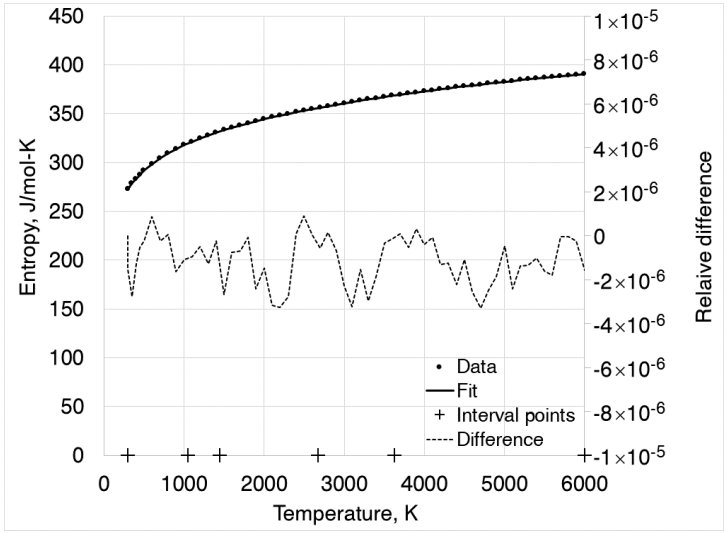


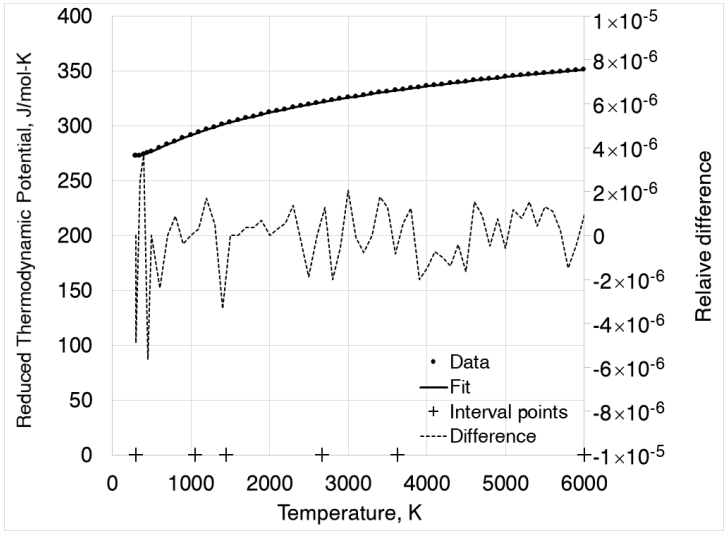
Lead Bromide (PbBr) – Br1Pb1(g)

MW = 287.1 g/mol – $\Delta_f H^o$(298.15 K) = 70.923 kJ/mol – S^o(298.15 K) = 272.471 J/mol-K – p^o = 0.1 MPa



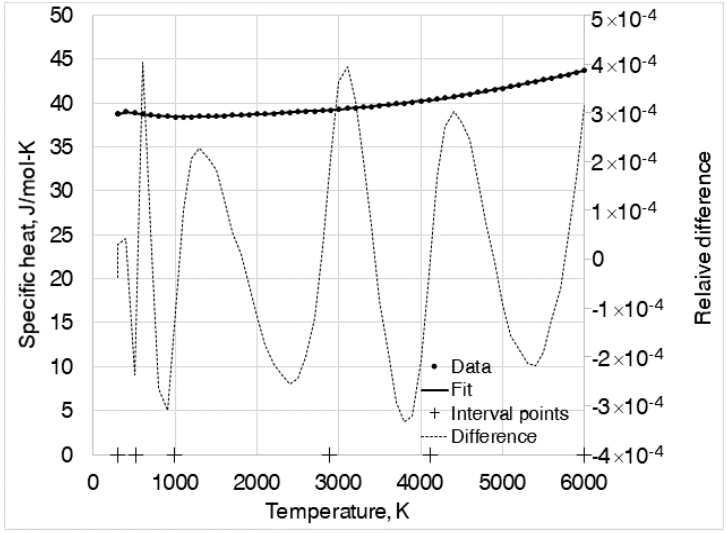


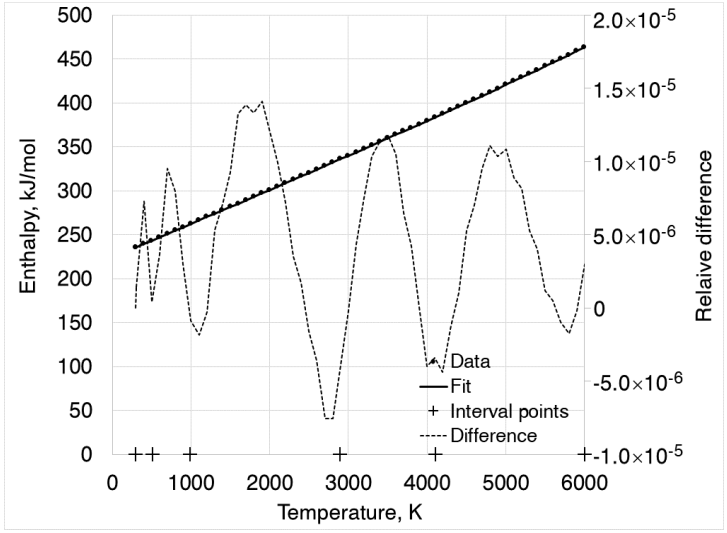


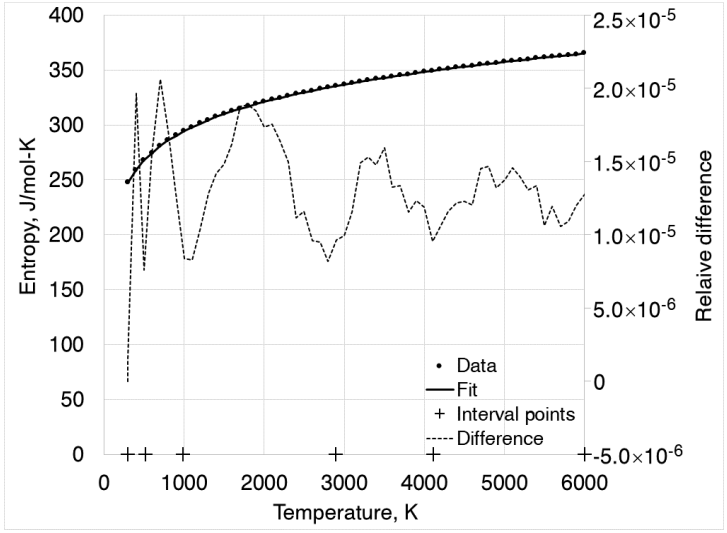


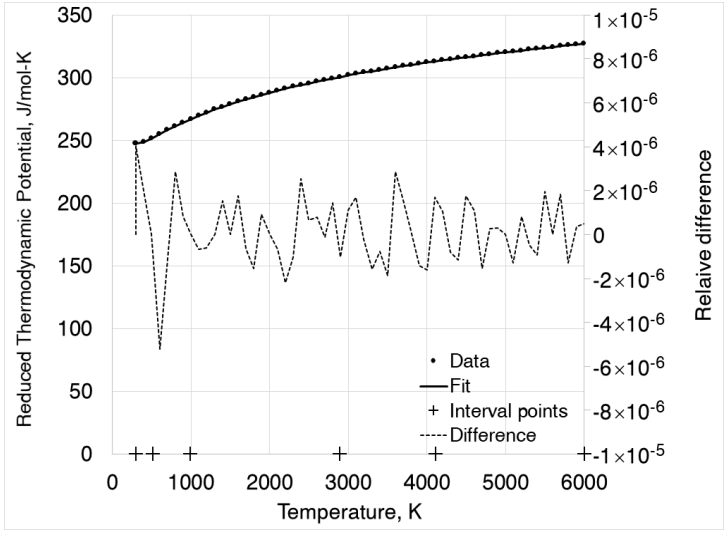
Bromosilylidyne (SiBr) – Br1Si1(g)

MW = 107.989 g/mol – $\Delta_f H^o$(298.15 K) = 235.325 kJ/mol – S^o(298.15 K) = 247.418 J/mol-K – p^o = 0.1 MPa



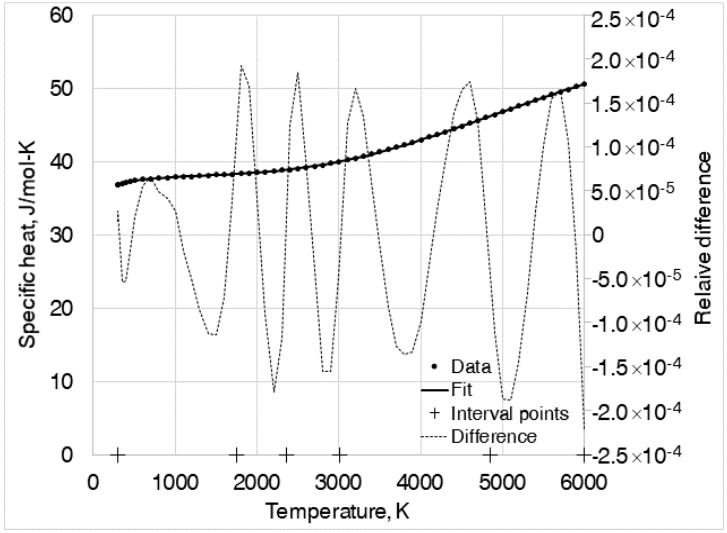


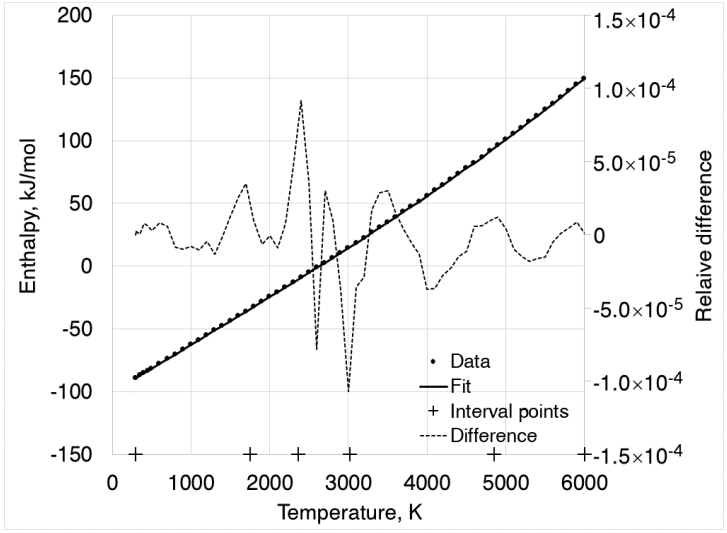


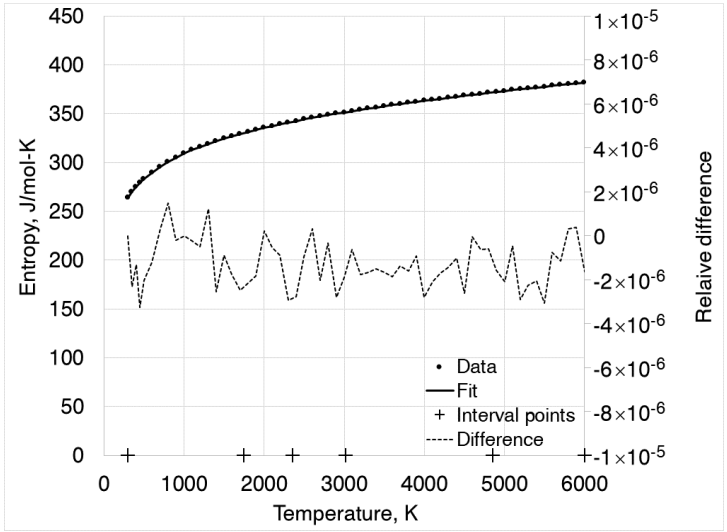


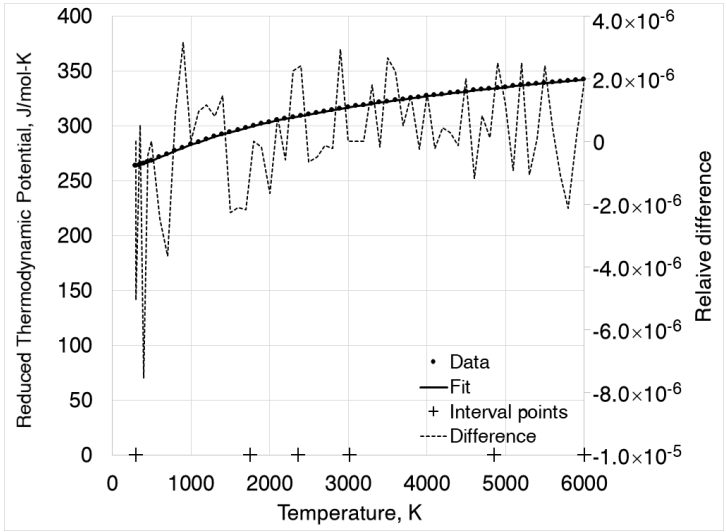
Strontium Bromide (SrBr) – Br1Sr1(g)

MW = 167.52 g/mol – $\Delta_f H^o$(298.15 K) = -89.119 kJ/mol – S^o(298.15 K) = 263.751 J/mol-K – p^o = 0.1 MPa



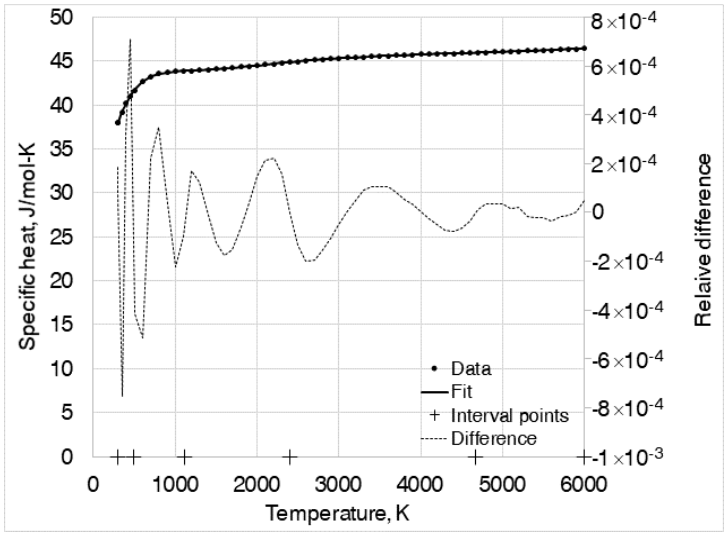


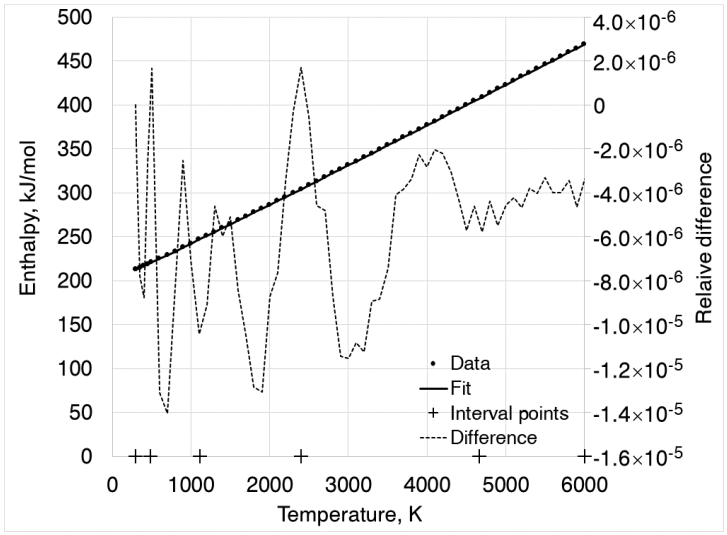


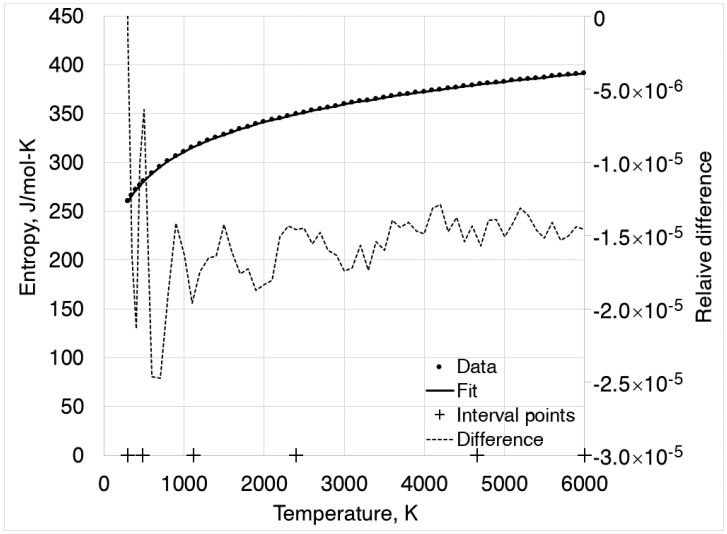


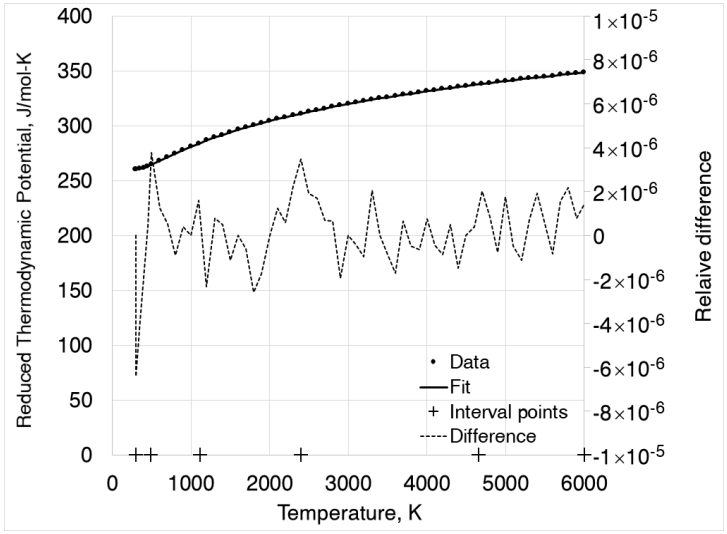
Titanium Bromide (TiBr) – Br1Ti1(g)

MW = 127.78 g/mol – $\Delta_f H^o$(298.15 K) = 212.547 kJ/mol – S^o(298.15 K) = 260.249 J/mol-K – p^o = 0.1 MPa



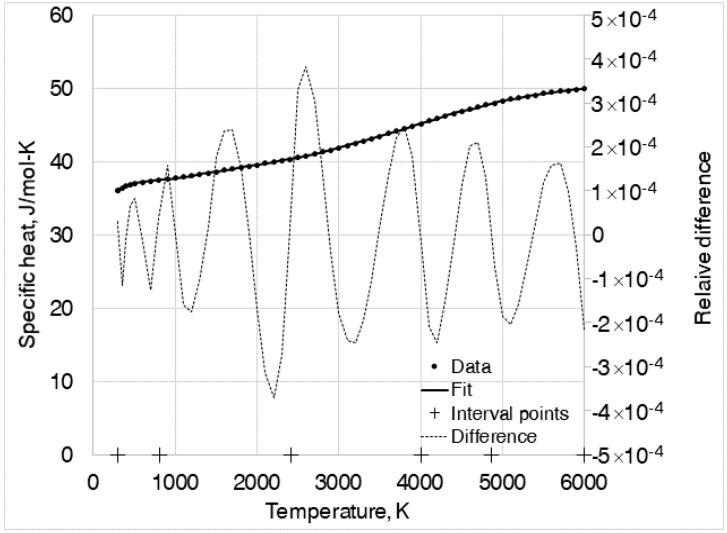


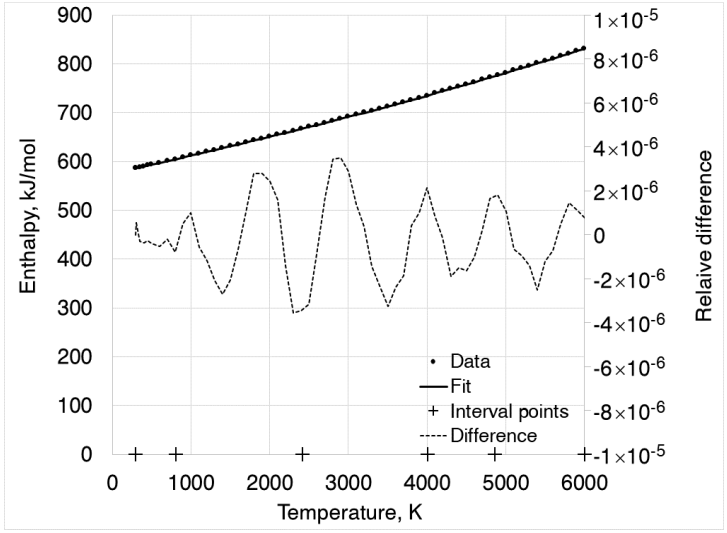


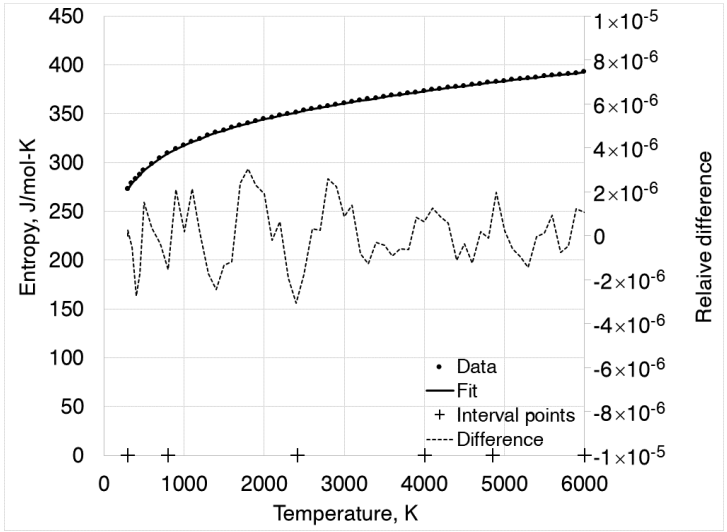


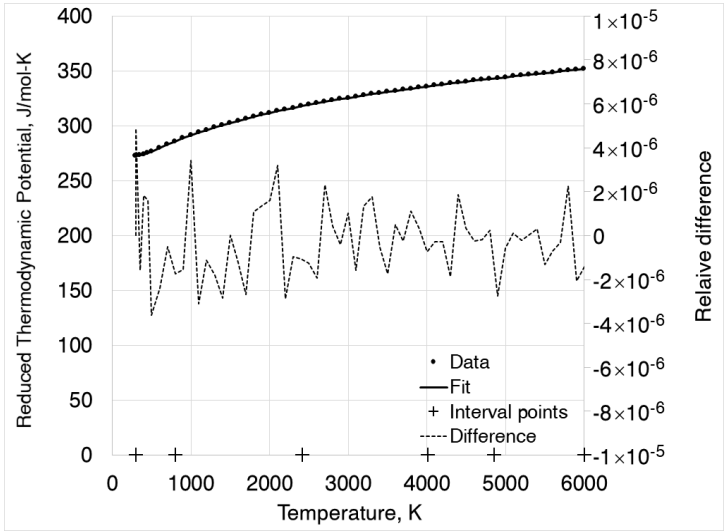
Tungsten Bromide (WBr) – Br1W1(g)

MW = 263.75 g/mol – $\Delta_f H^o$(298.15 K) = 586.178 kJ/mol – S^o(298.15 K) = 272.617 J/mol-K – p^o = 0.1 MPa



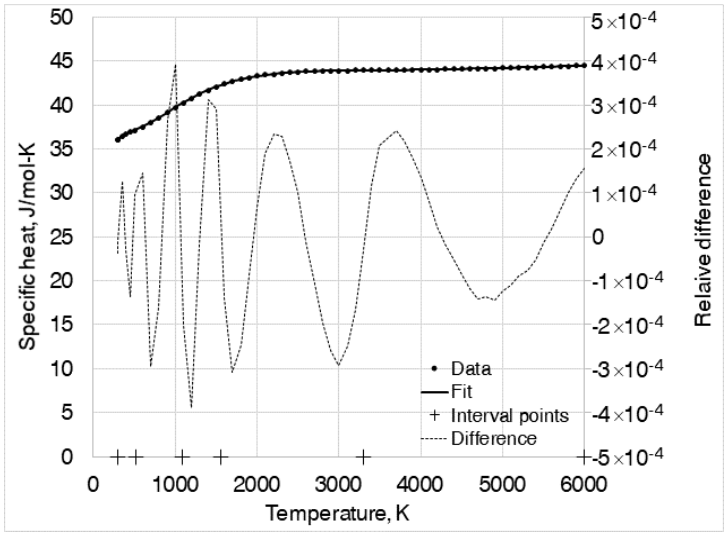


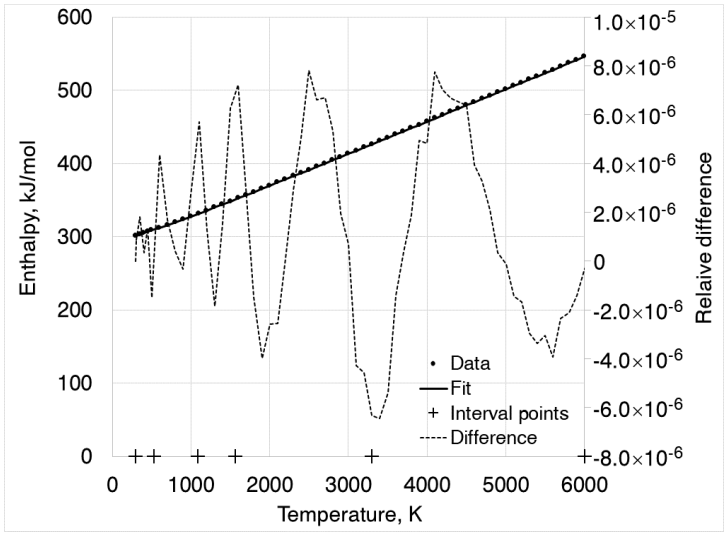


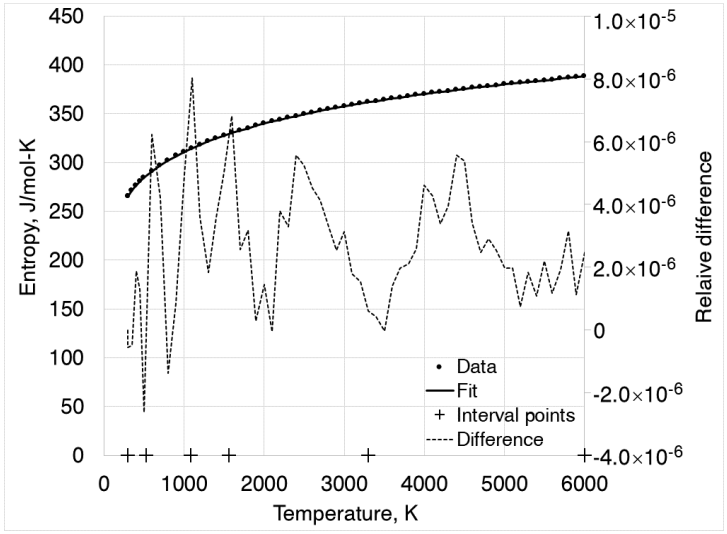


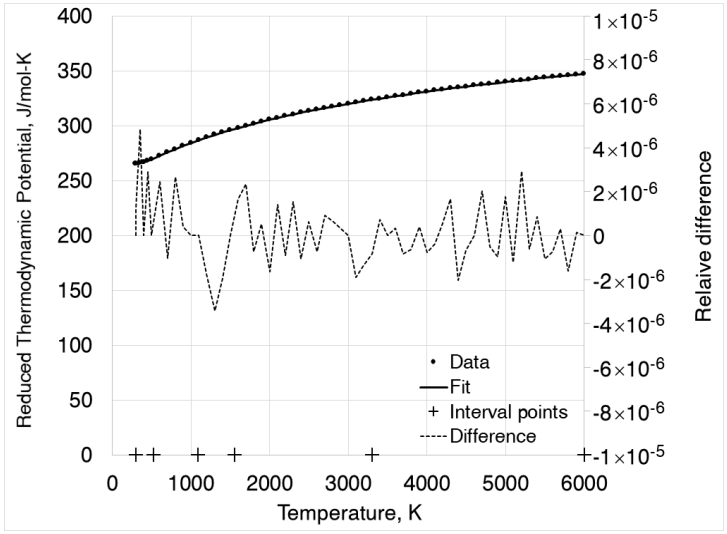
Zirconium Bromide (ZrBr) – Br1Zr1(g)

MW = 171.12 g/mol – $\Delta_f H^o$(298.15 K) = 301.248 kJ/mol – S^o(298.15 K) = 265.557 J/mol-K – p^o = 0.1 MPa



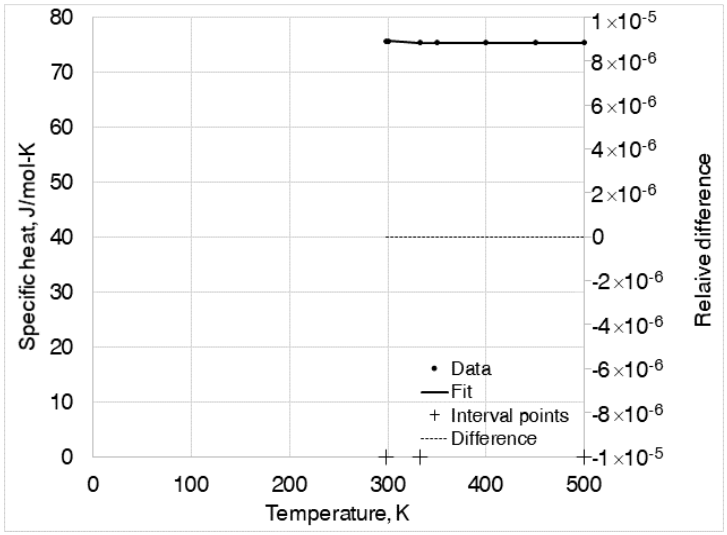


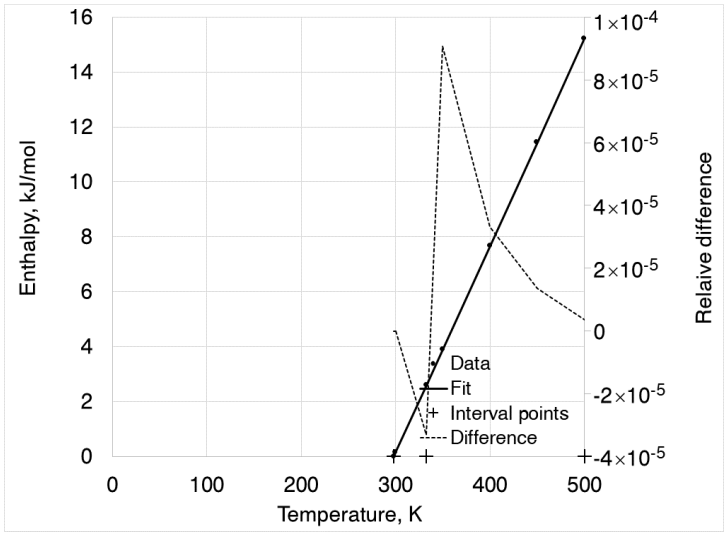


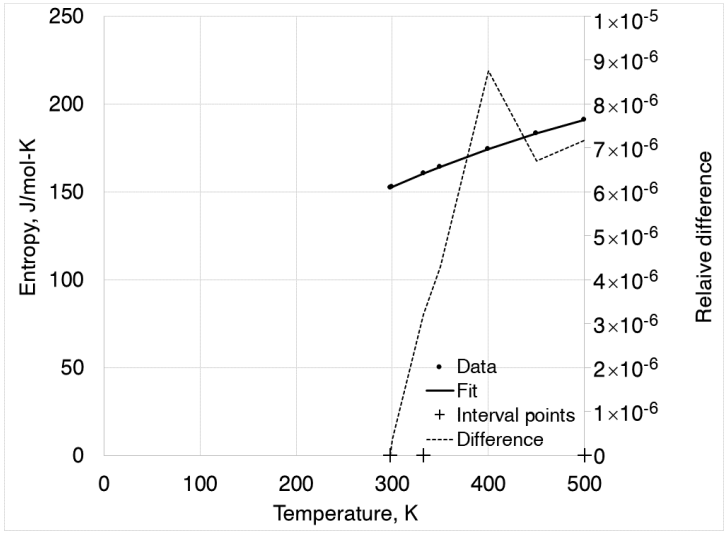


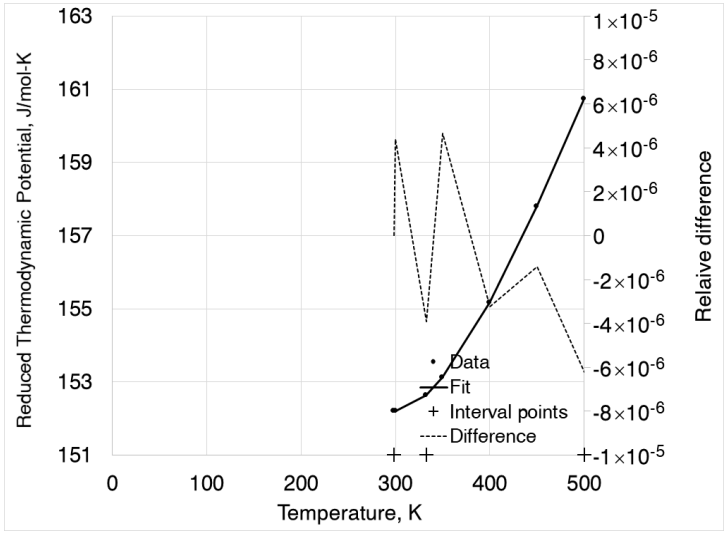
Bromine (Br_2) – Br2(cr,l)

MW = 159.808 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 152.206 J/mol-K – p^o = 0.1 MPa



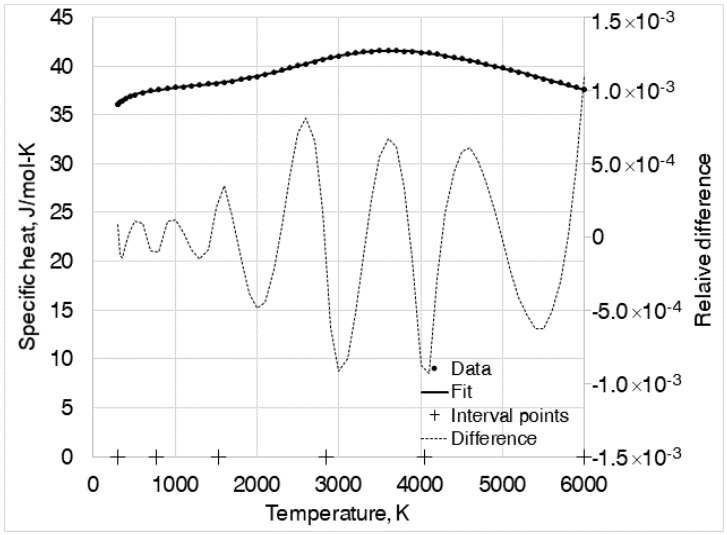


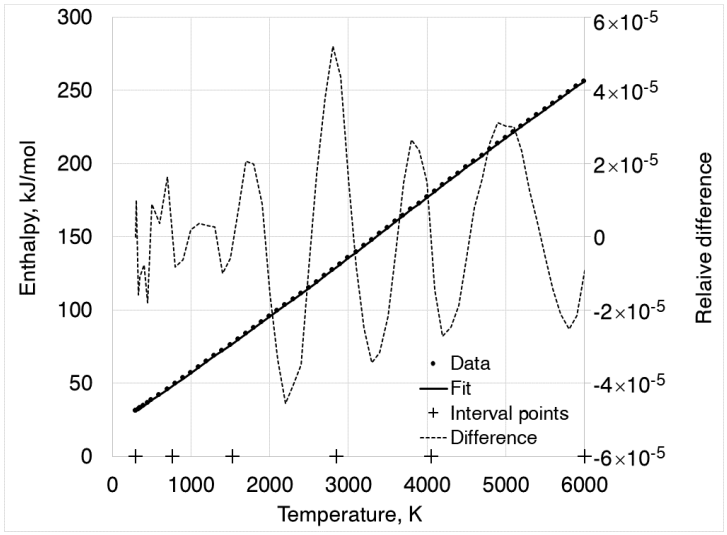


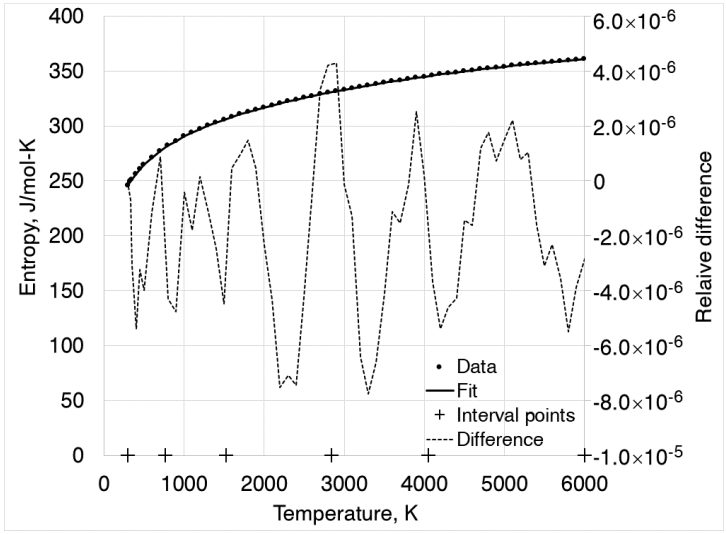


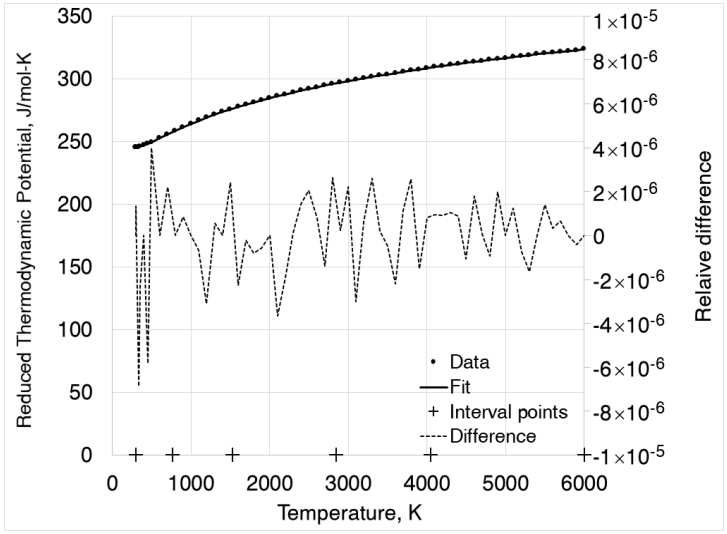
Bromine (Br_2) – Br2(g)

MW = 159.808 g/mol – $\Delta_f H^o$(298.15 K) = 30.91 kJ/mol – S^o(298.15 K) = 245.394 J/mol-K – p^o = 0.1 MPa



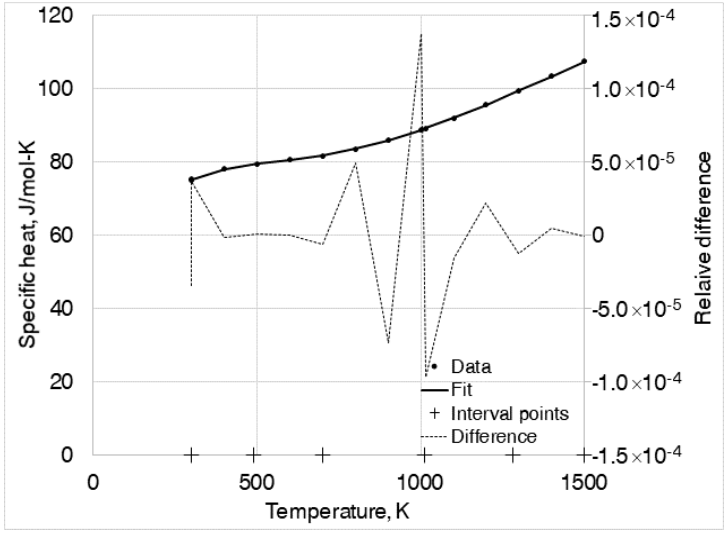


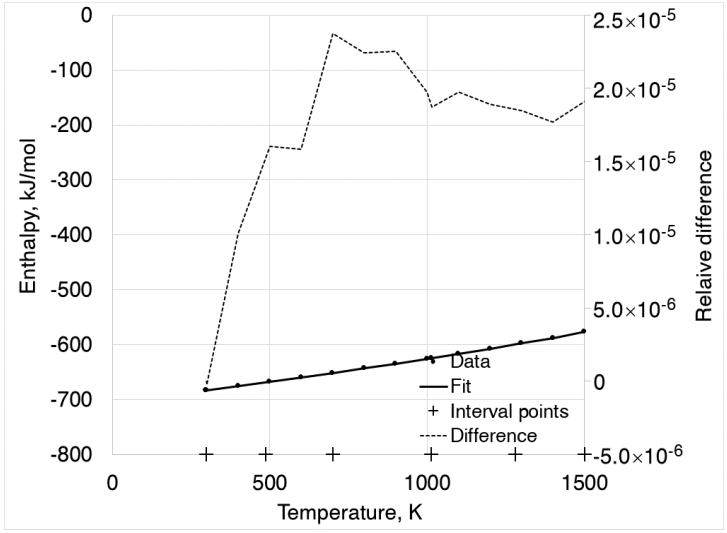


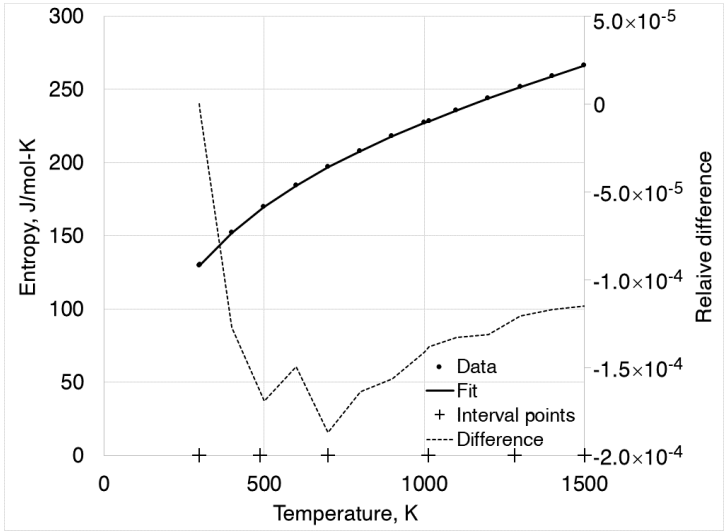


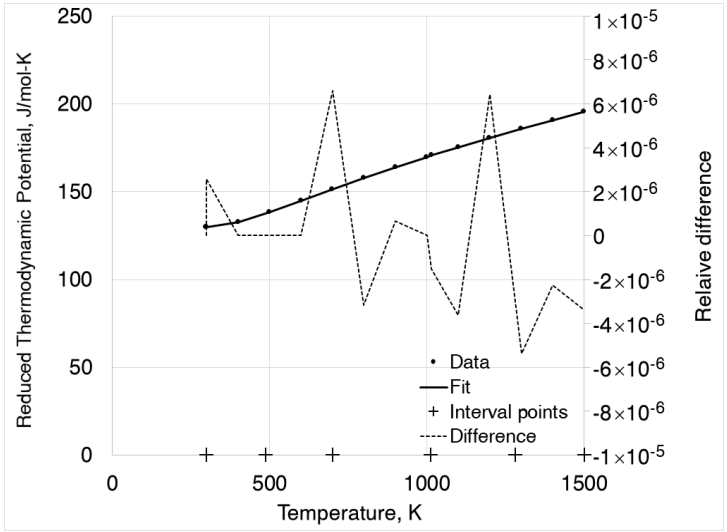
Calcium Bromide ($CaBr_2$) – Br2Ca1(cr)

MW = 199.89 g/mol – $\Delta_f H^o$(298.15 K) = -683.247 kJ/mol – S^o(298.15 K) = 129.704 J/mol-K – p^o = 0.1 MPa



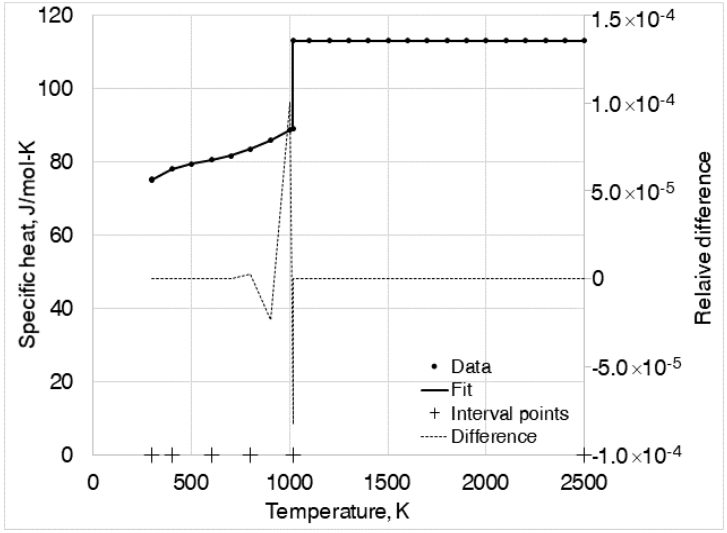


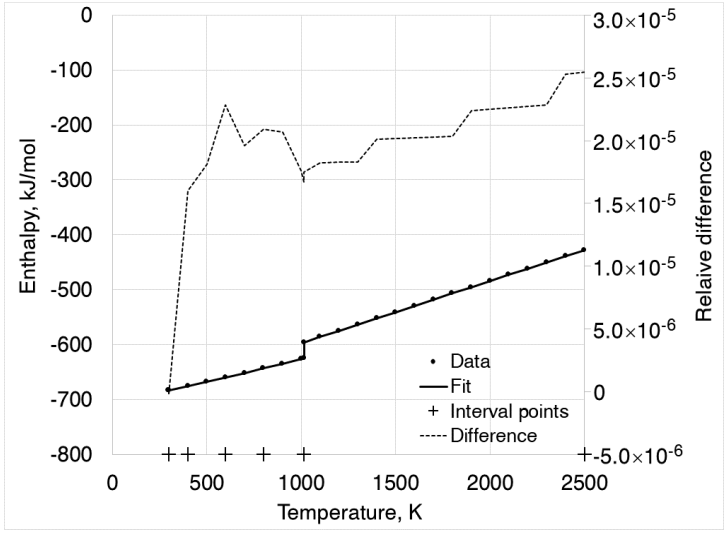


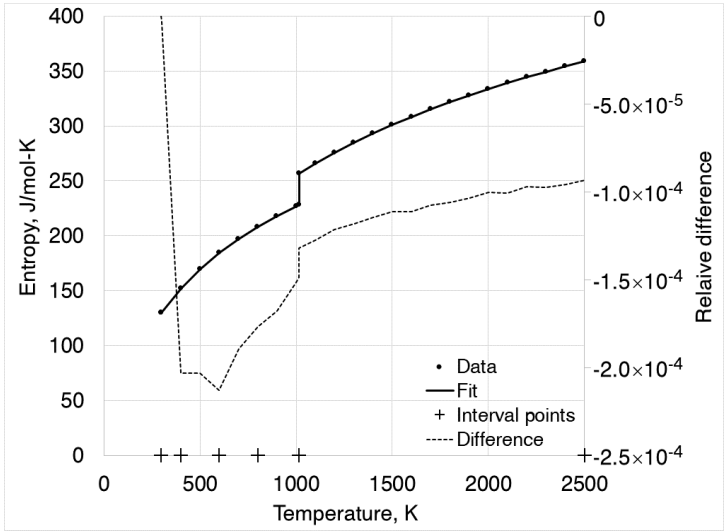


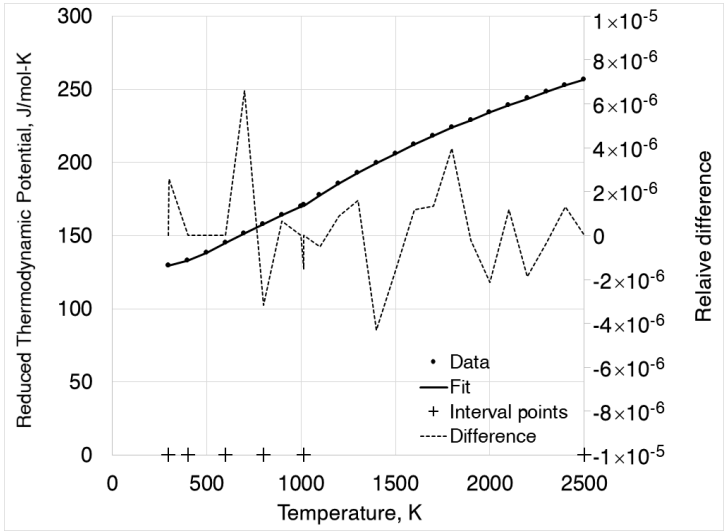
Calcium Bromide ($CaBr_2$) – Br2Ca1(cr,l)

MW = 199.89 g/mol – $\Delta_f H^o$(298.15 K) = -683.247 kJ/mol – S^o(298.15 K) = 129.704 J/mol-K – p^o = 0.1 MPa



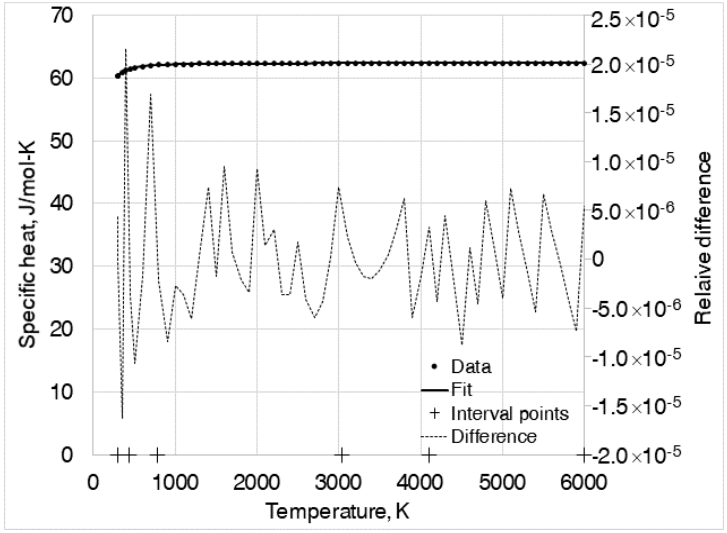


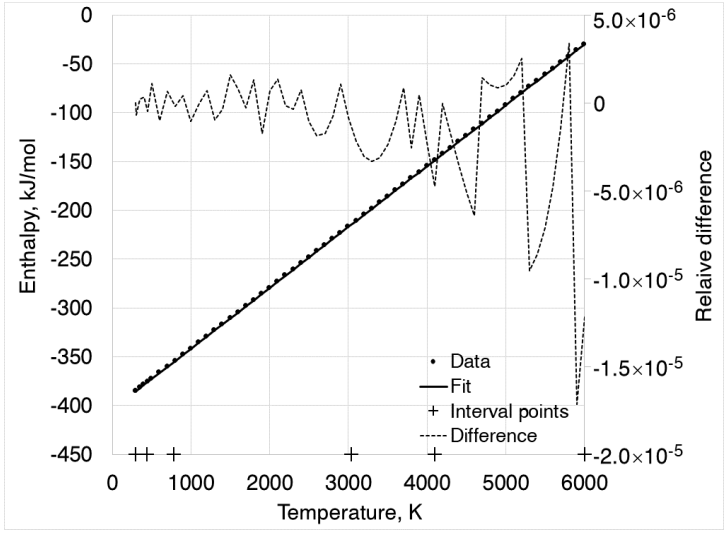


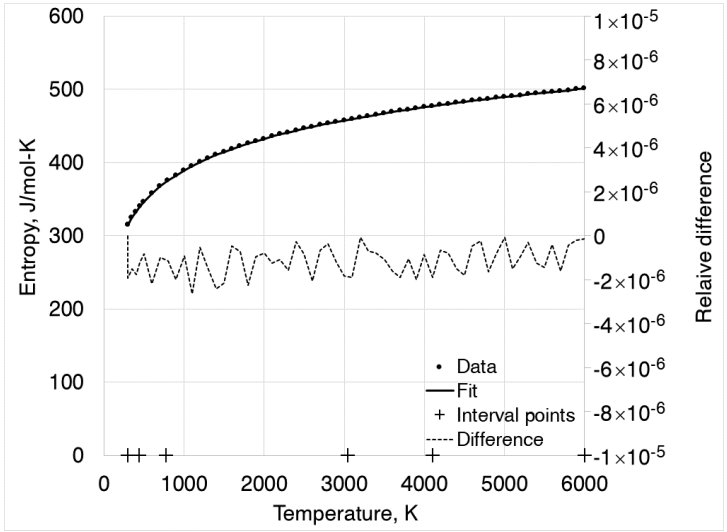


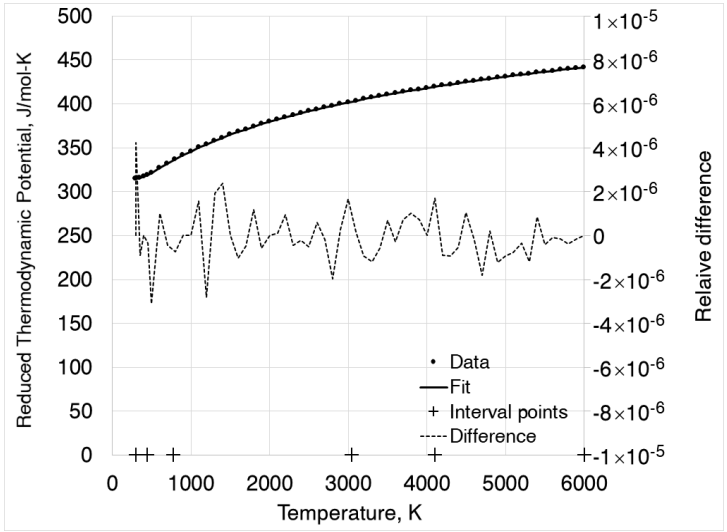
Calcium Bromide ($CaBr_2$) – Br2Ca1(g)

MW = 199.89 g/mol – $\Delta_f H^o$(298.15 K) = -384.928 kJ/mol – S^o(298.15 K) = 314.752 J/mol-K – p^o = 0.1 MPa



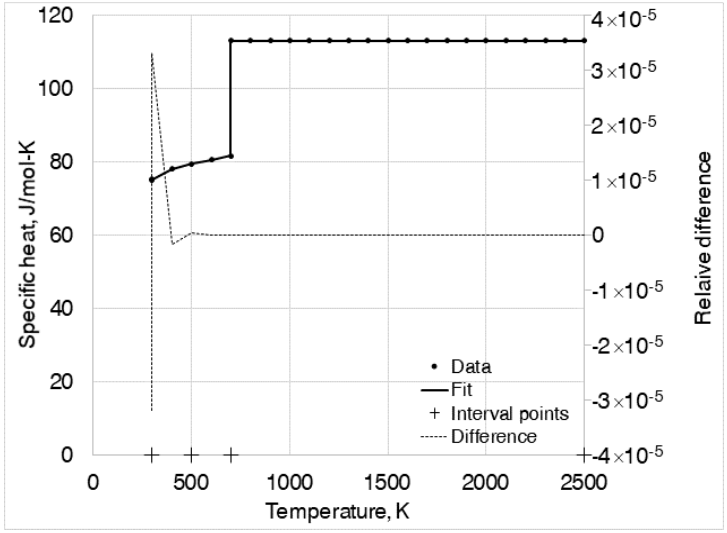


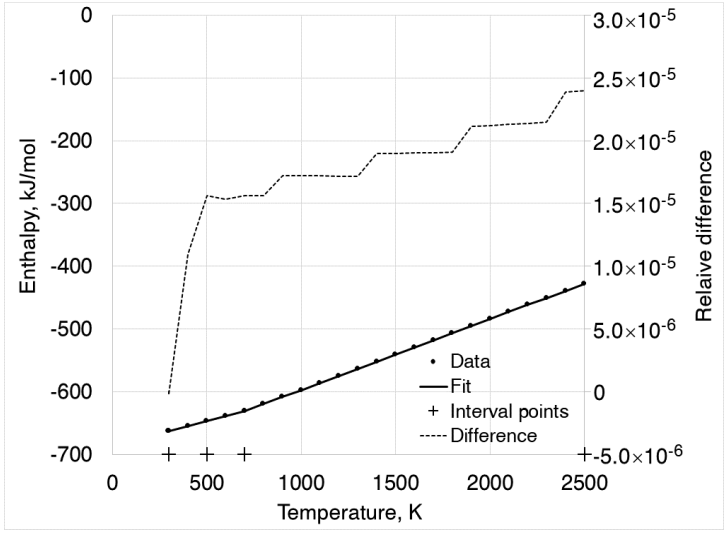


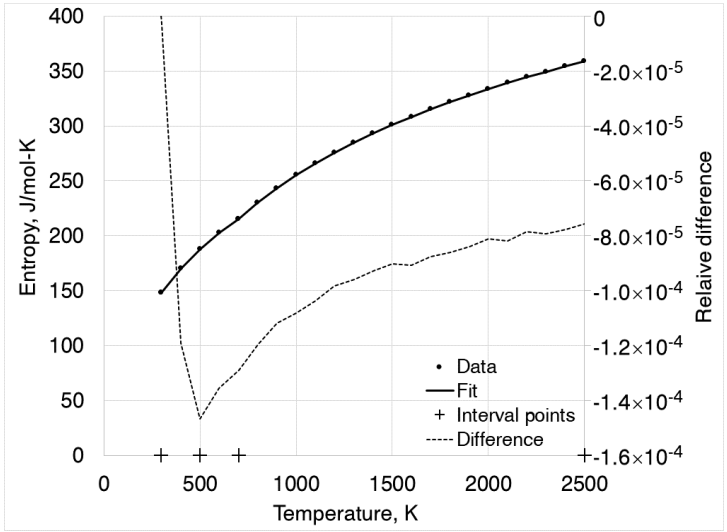


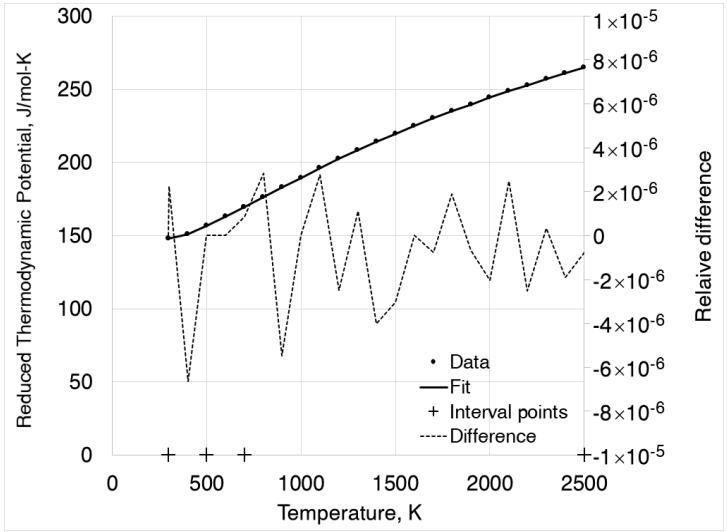
Calcium Bromide ($CaBr_2$) – Br2Ca1(l)

MW = 199.89 g/mol – $\Delta_f H^o$(298.15 K) = -662.989 kJ/mol – S^o(298.15 K) = 147.861 J/mol-K – p^o = 0.1 MPa



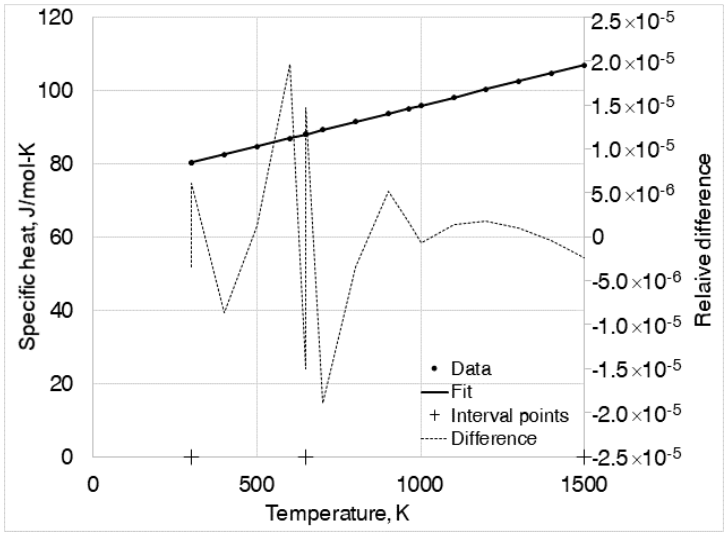


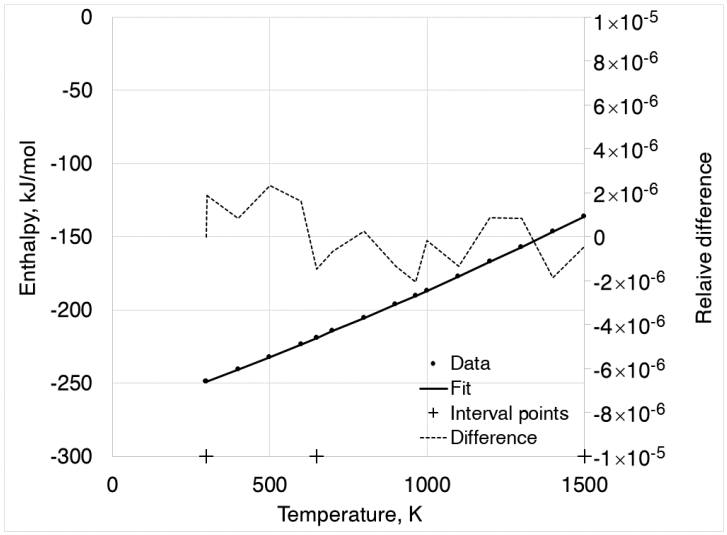


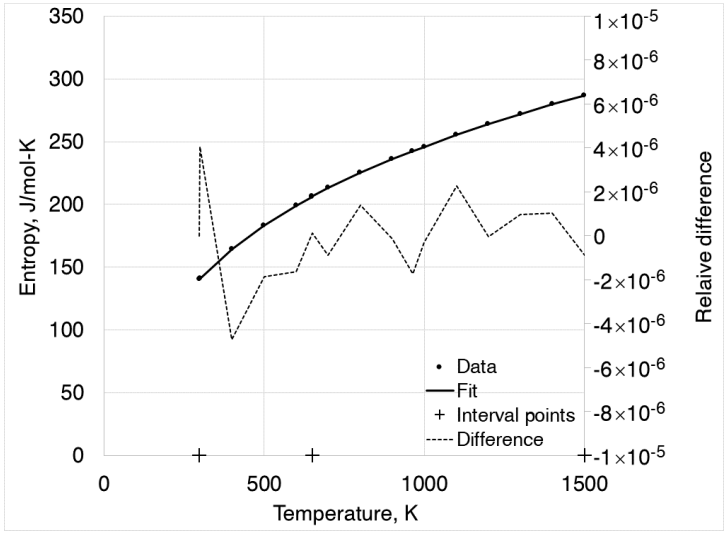


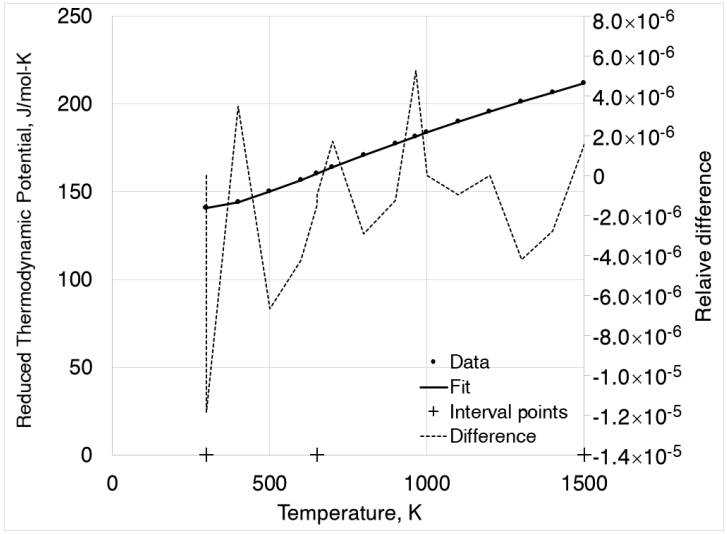
Iron Bromide ($FeBr_2$) – Br2Fe1(cr)

MW = 215.655 g/mol – $\Delta_f H^o$(298.15 K) = -248.948 kJ/mol – S^o(298.15 K) = 140.666 J/mol-K – p^o = 0.1 MPa



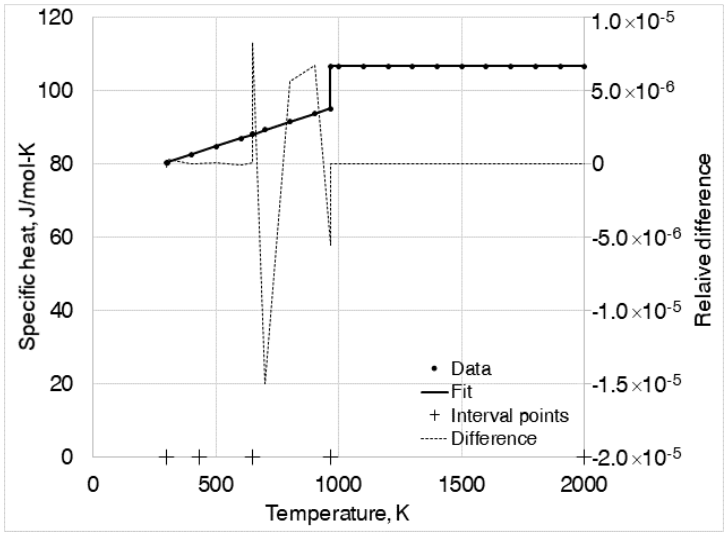


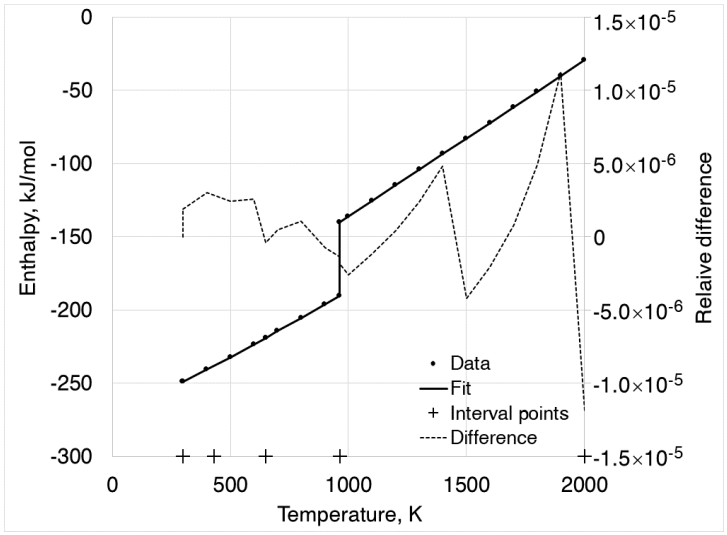


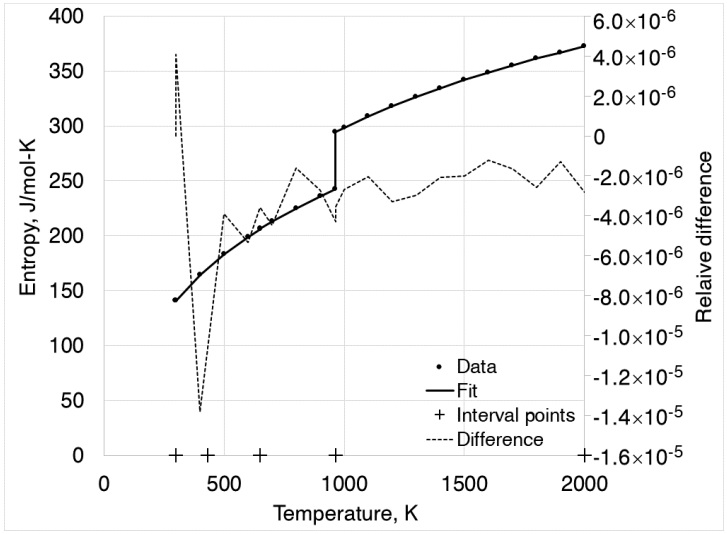


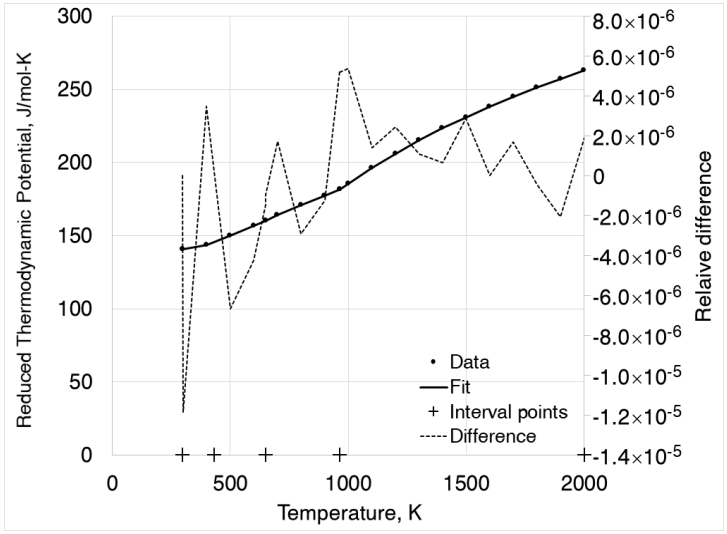
Iron Bromide ($FeBr_2$) – Br2Fe1(cr,l)

MW = 215.655 g/mol – $\Delta_f H^o$(298.15 K) = -248.948 kJ/mol – S^o(298.15 K) = 140.666 J/mol-K – p^o = 0.1 MPa



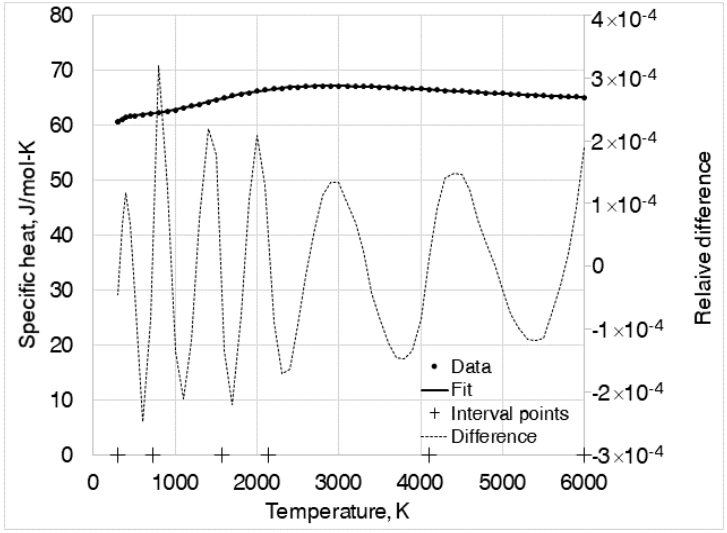


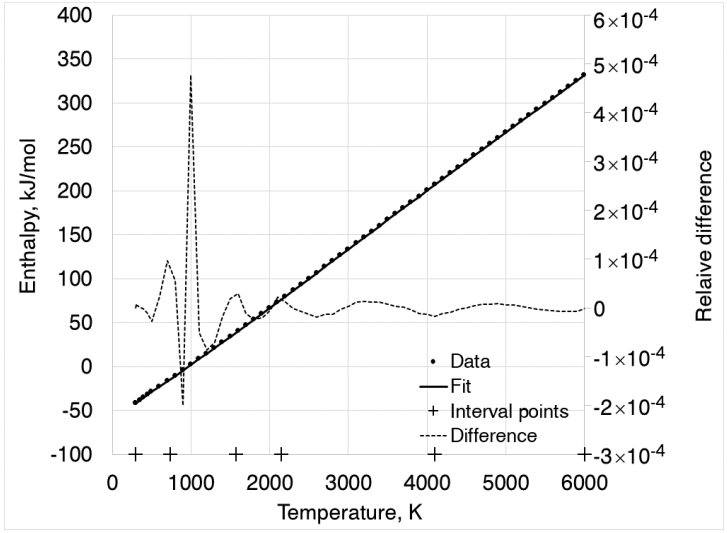


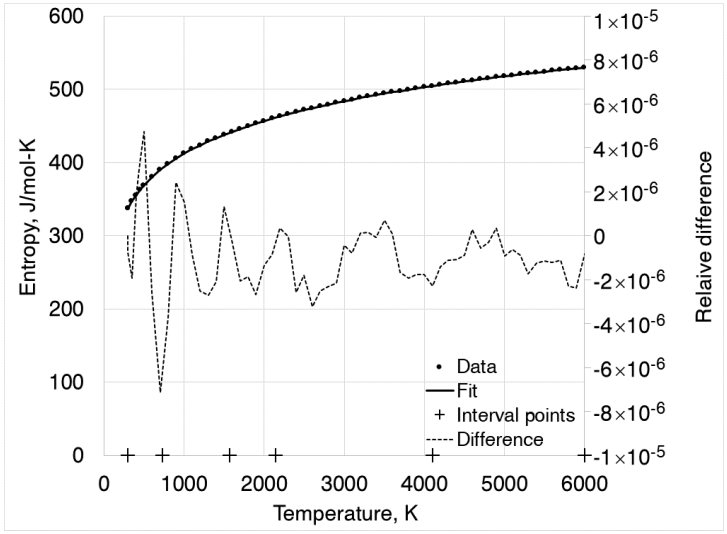


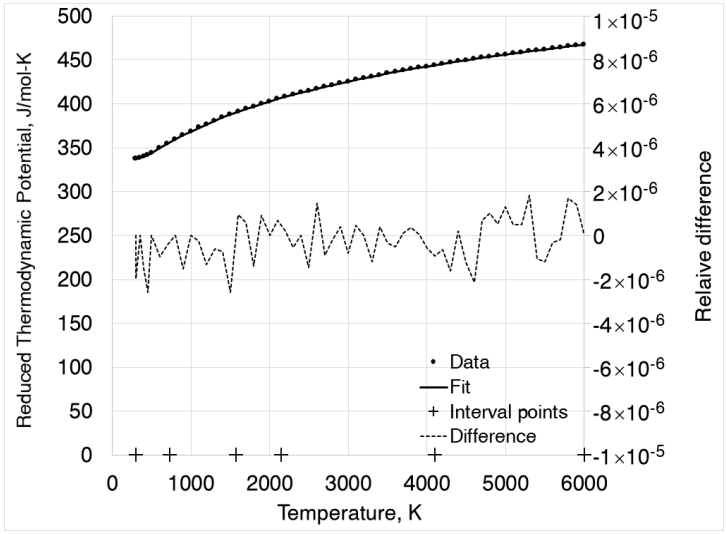
Iron Bromide ($FeBr_2$) – Br2Fe1(g)

MW = 215.655 g/mol – $\Delta_f H^o$(298.15 K) = -41.422 kJ/mol – S^o(298.15 K) = 337.359 J/mol-K – p^o = 0.1 MPa



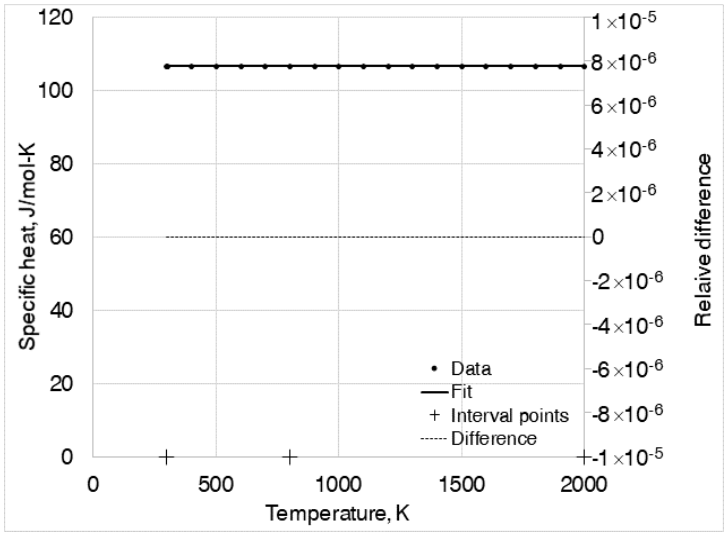


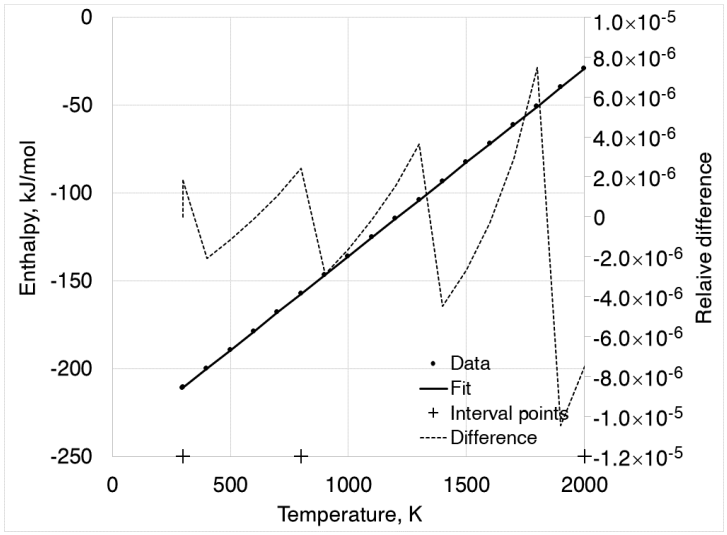


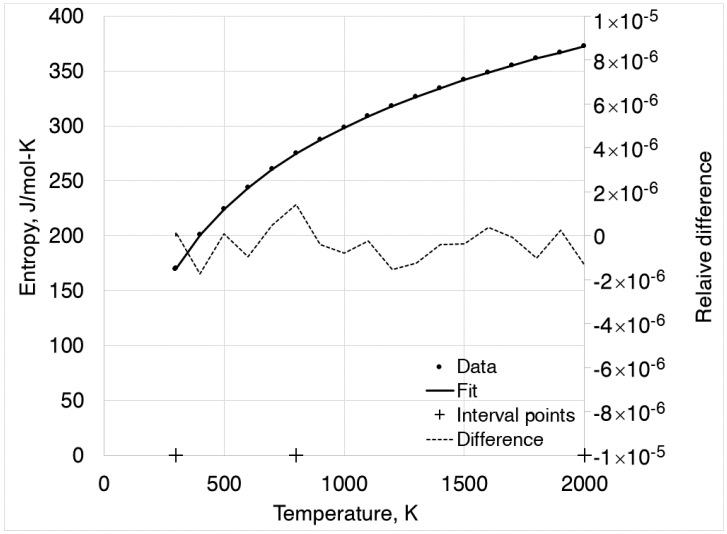


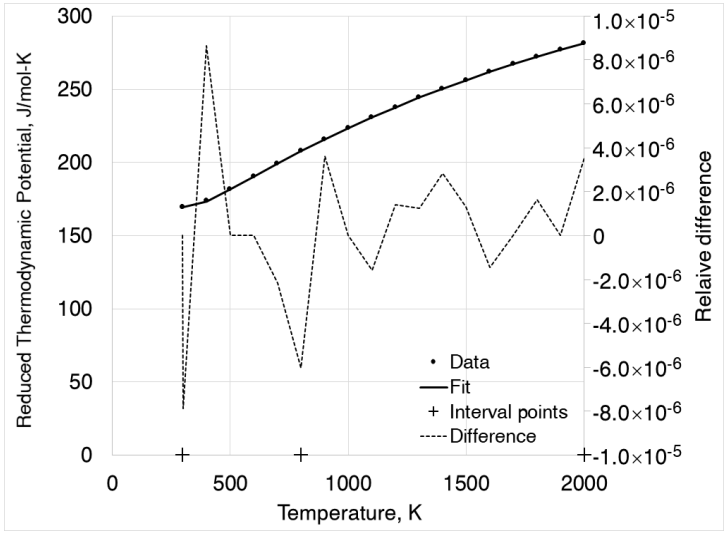
Iron Bromide ($FeBr_2$) – Br2Fe1(l)

MW = 215.655 g/mol – $\Delta_f H^o$(298.15 K) = -211.005 kJ/mol – S^o(298.15 K) = 169.376 J/mol-K – p^o = 0.1 MPa



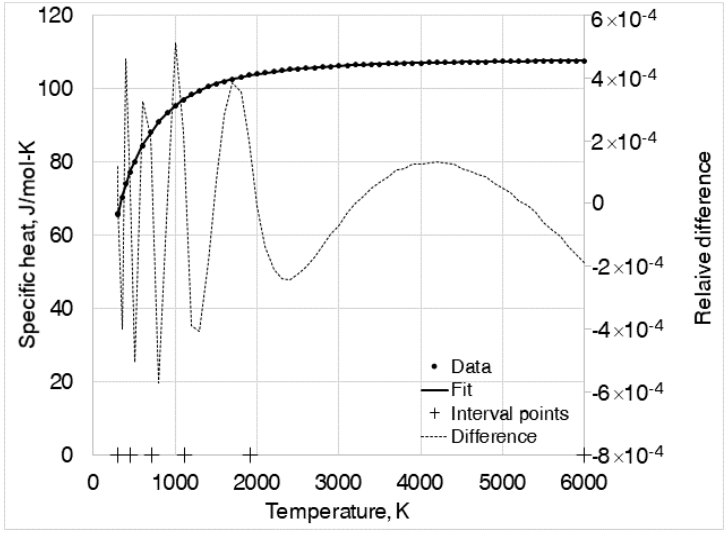


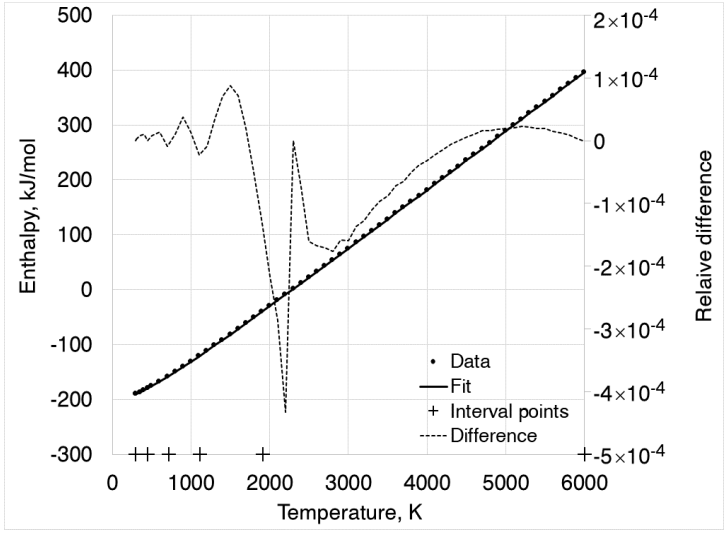


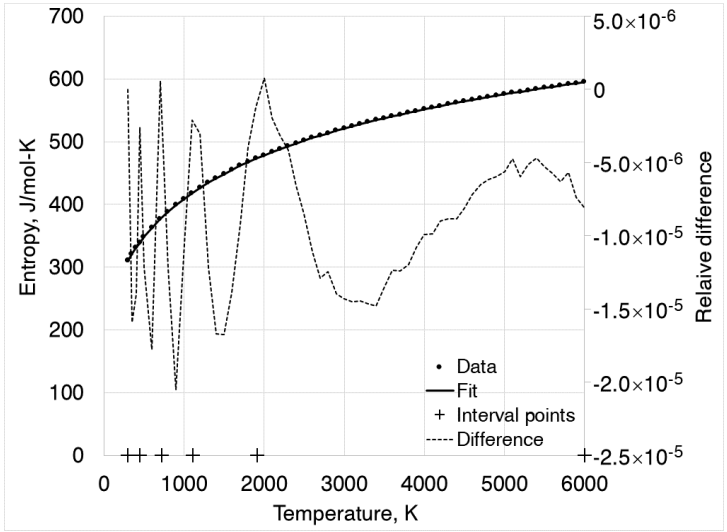


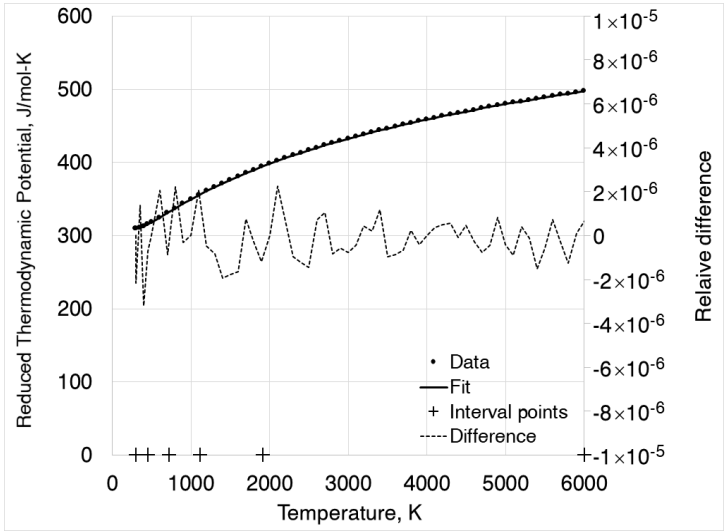
Dibromosilane (SiH_2Br_2) – Br2H2Si1(g)

MW = 189.909 g/mol – $\Delta_f H^o$(298.15 K) = -190.372 kJ/mol – S^o(298.15 K) = 310.051 J/mol-K – p^o = 0.1 MPa



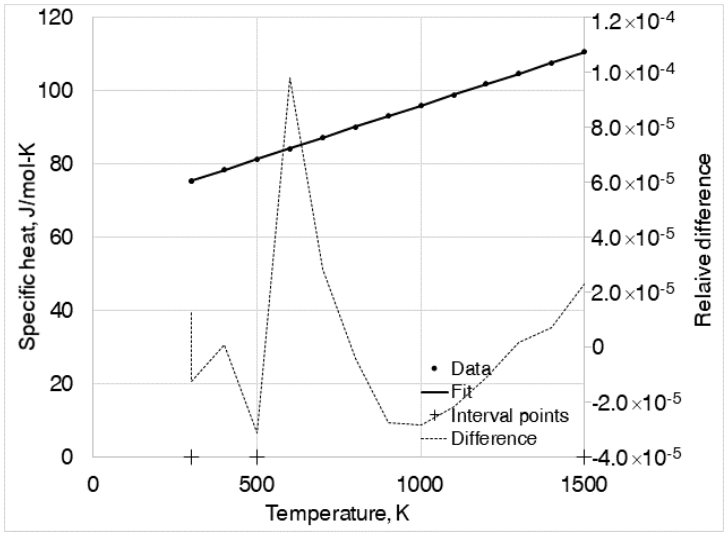


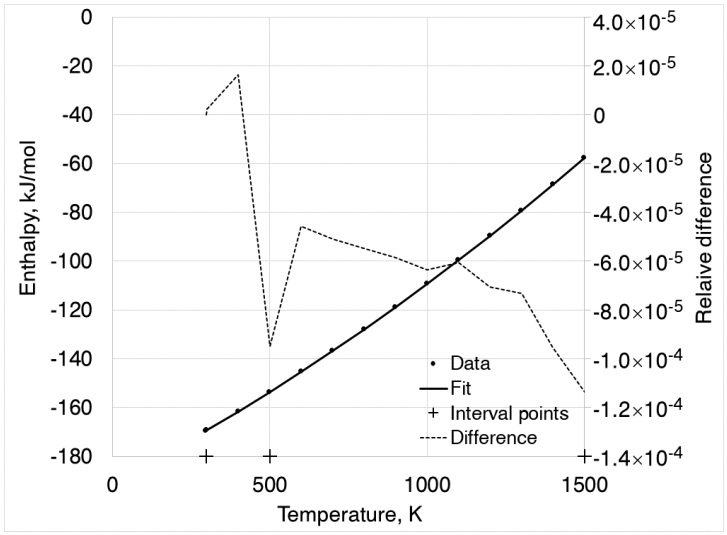


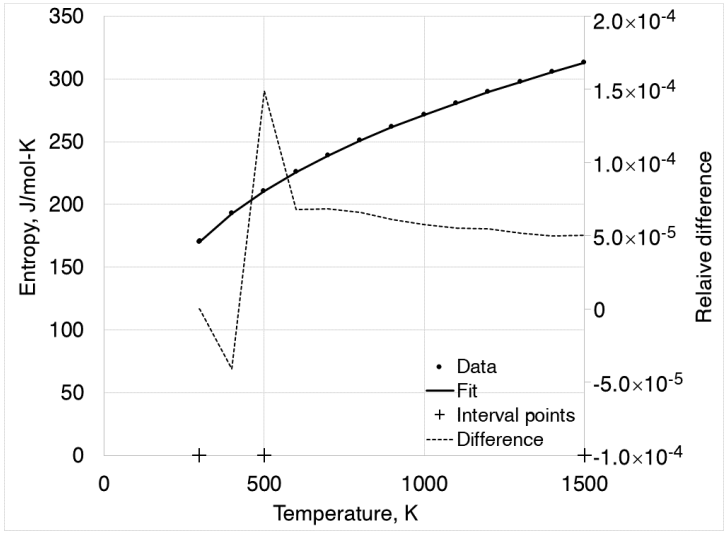


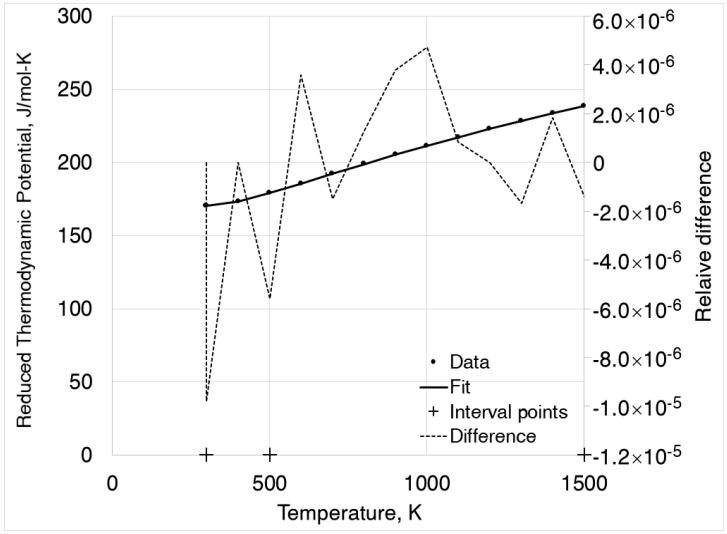
Mercury Bromide ($HgBr_2$) – Br2Hg1(cr)

MW = 360.4 g/mol – $\Delta_f H^o$(298.15 K) = -169.452 kJ/mol – S^o(298.15 K) = 170.314 J/mol-K – p^o = 0.1 MPa



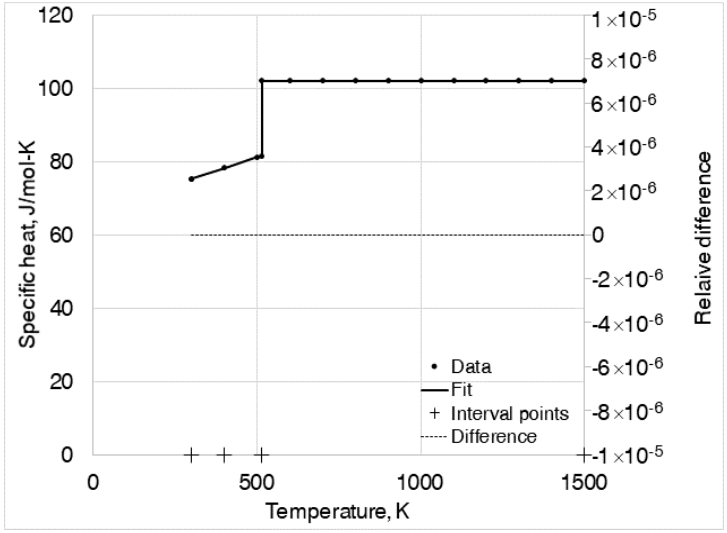


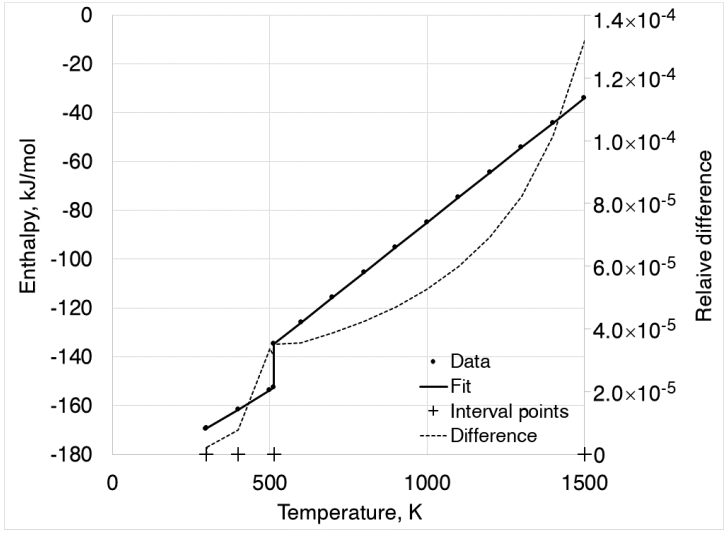


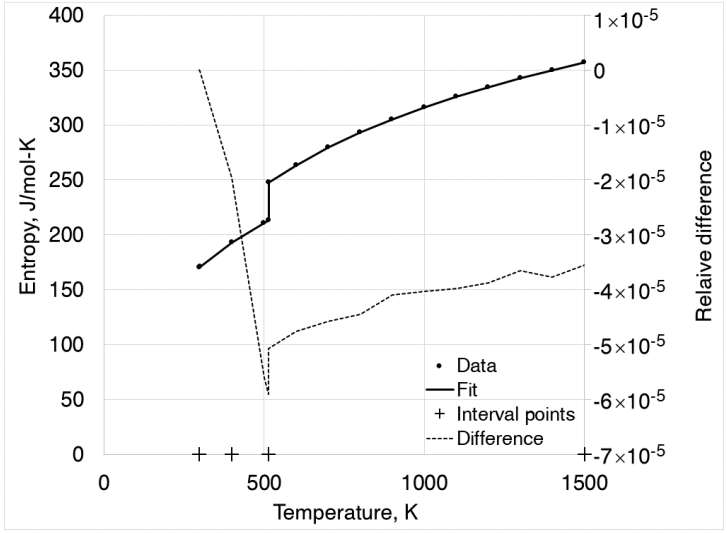


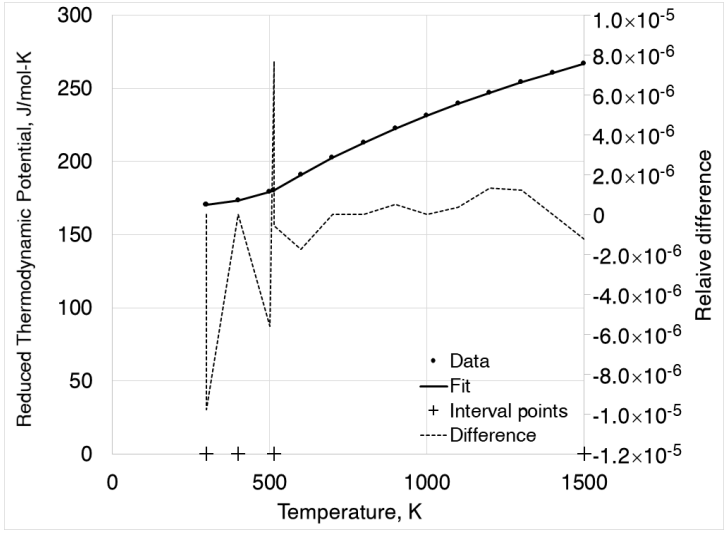
Mercury Bromide ($HgBr_2$) – Br2Hg1(cr,l)

MW = 360.4 g/mol – $\Delta_f H^o$(298.15 K) = -169.452 kJ/mol – S^o(298.15 K) = 170.314 J/mol-K – p^o = 0.1 MPa



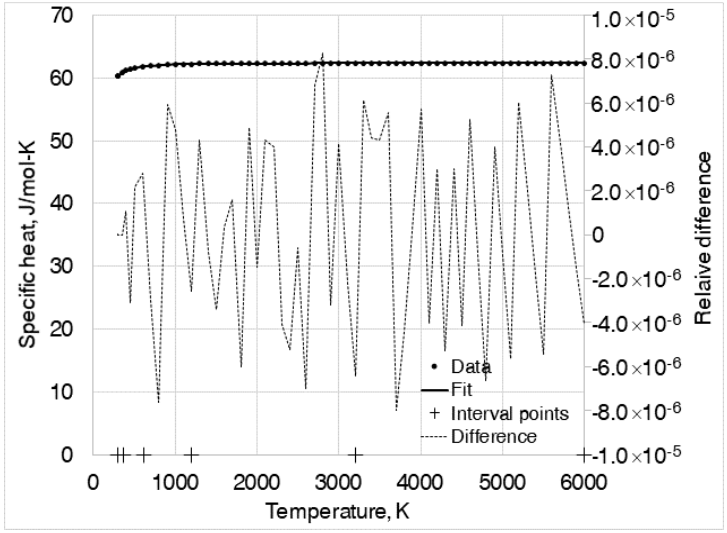


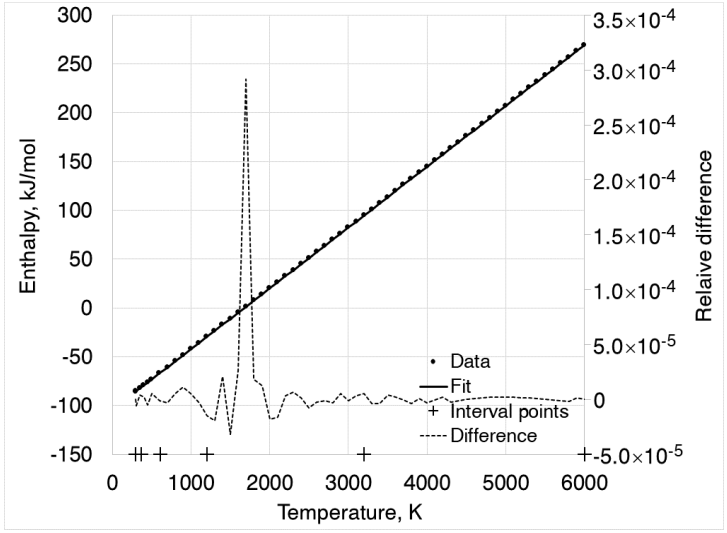


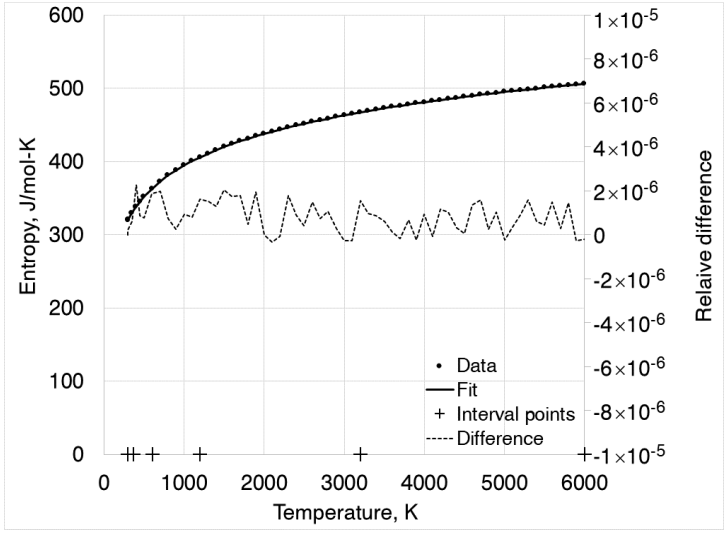


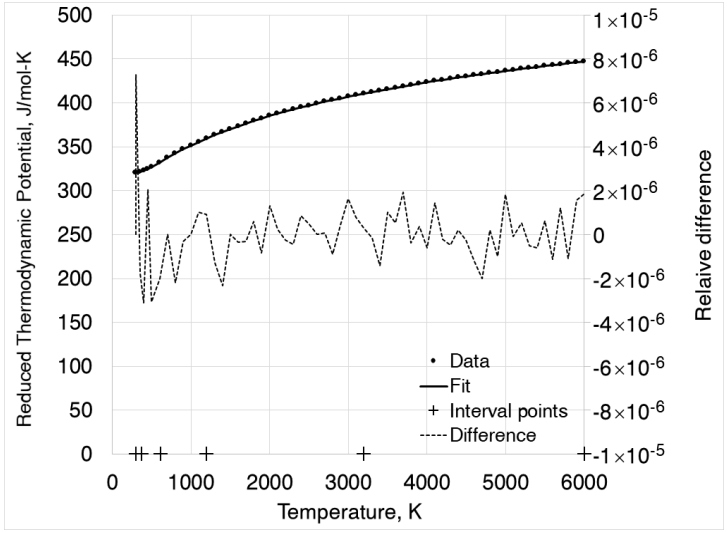
Mercury Bromide ($HgBr_2$) – Br2Hg1(g)

MW = 360.4 g/mol – $\Delta_f H^o$(298.15 K) = -85.454 kJ/mol – S^o(298.15 K) = 320.223 J/mol-K – p^o = 0.1 MPa



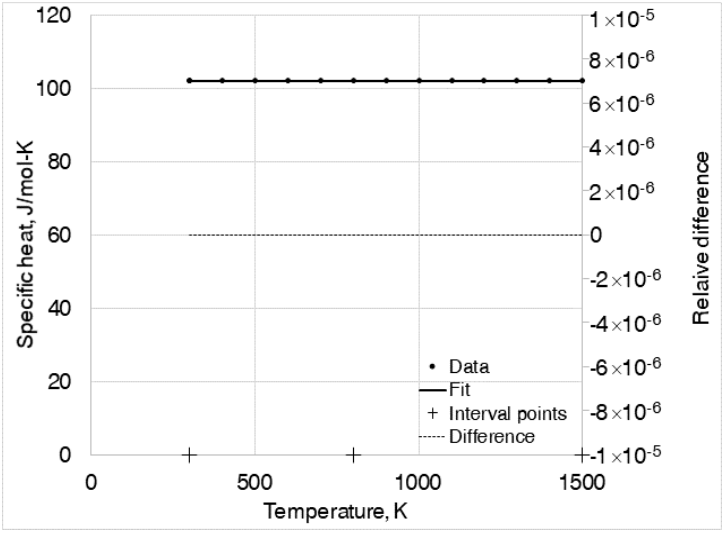


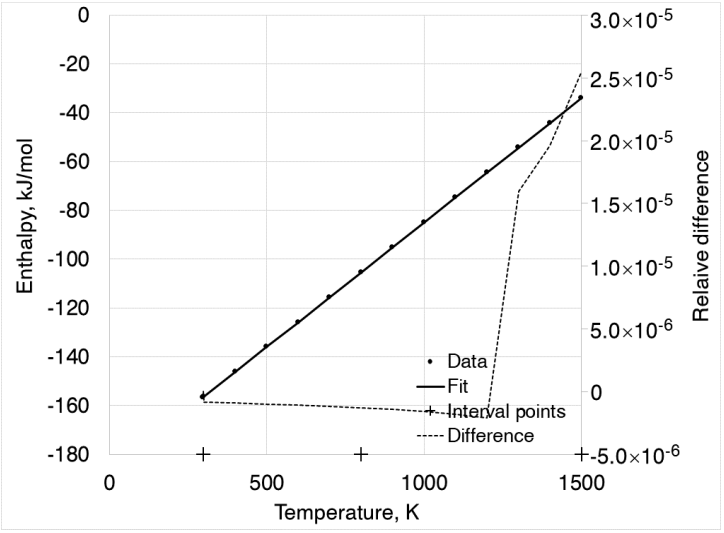


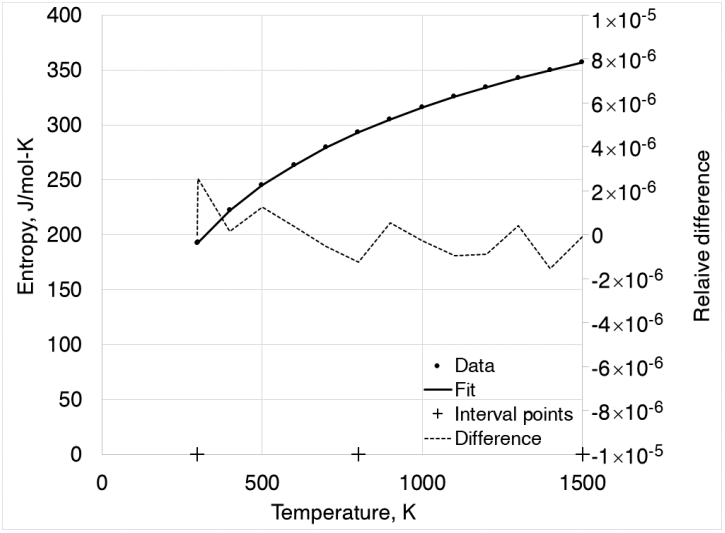


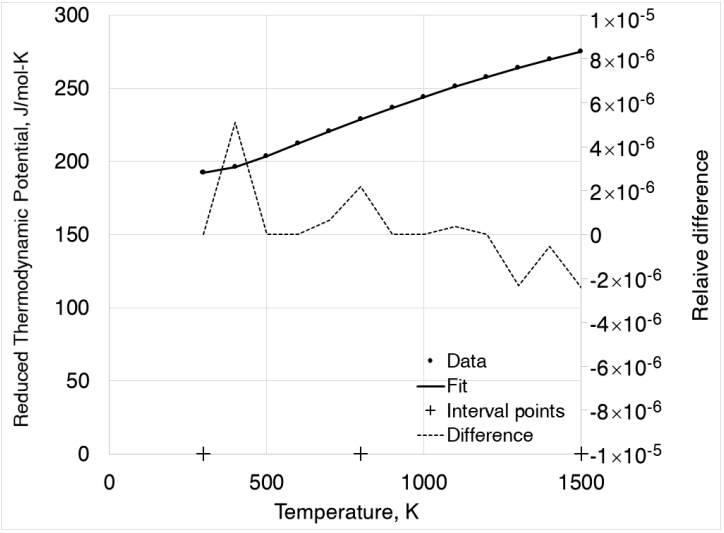
Mercury Bromide ($HgBr_2$) – Br2Hg1(l)

MW = 360.4 g/mol – $\Delta_f H^o$(298.15 K) = -156.626 kJ/mol – S^o(298.15 K) = 192.172 J/mol-K – p^o = 0.1 MPa



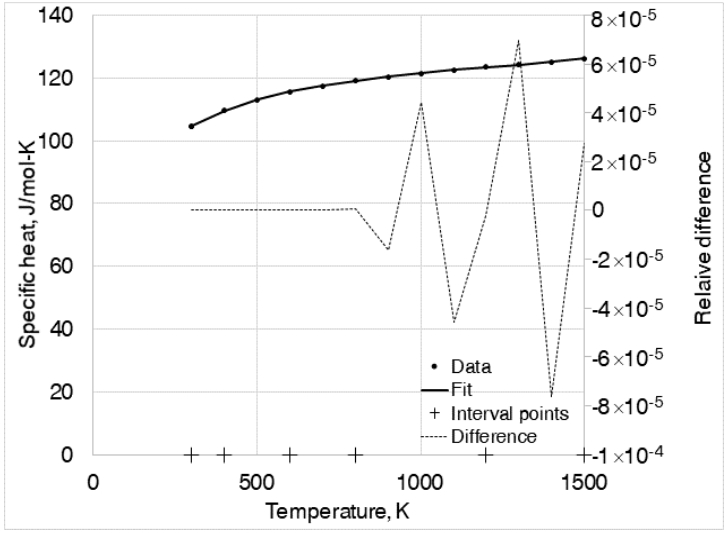


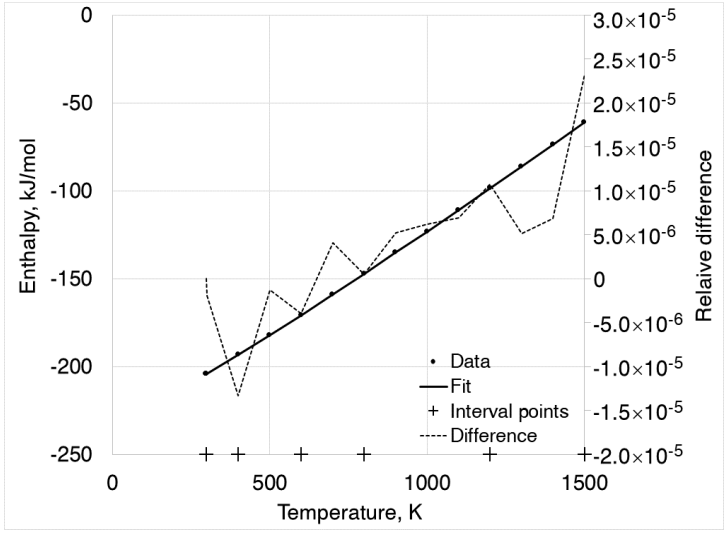


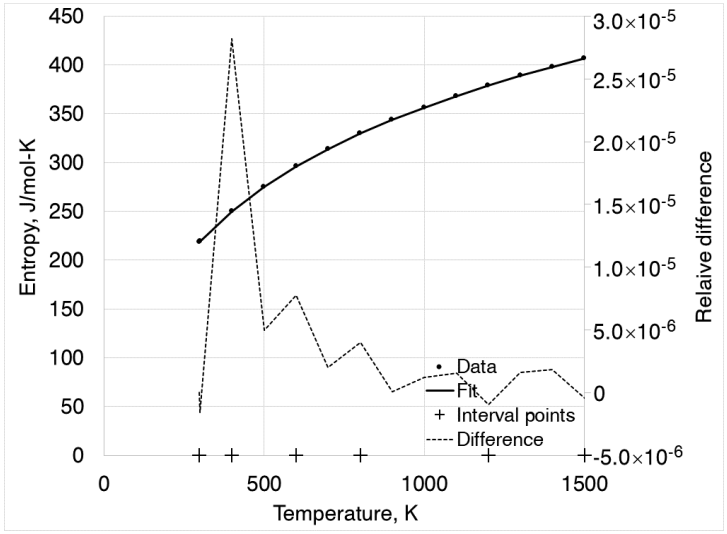


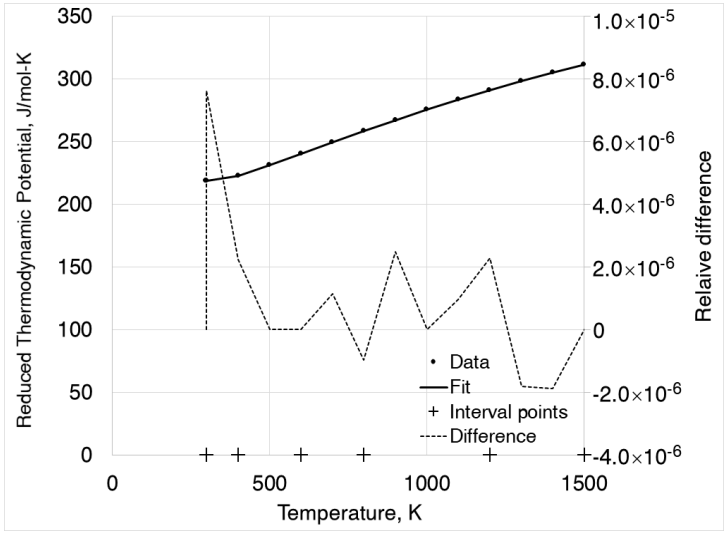
Mercury Bromide (Hg_2Br_2) – Br2Hg2(cr)

MW = 560.99 g/mol – $\Delta_f H^o$(298.15 K) = -204.179 kJ/mol – S^o(298.15 K) = 218.752 J/mol-K – p^o = 0.1 MPa



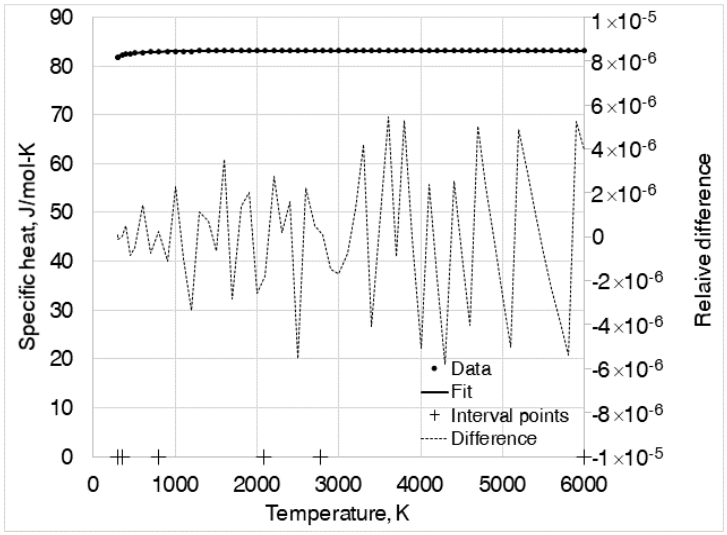


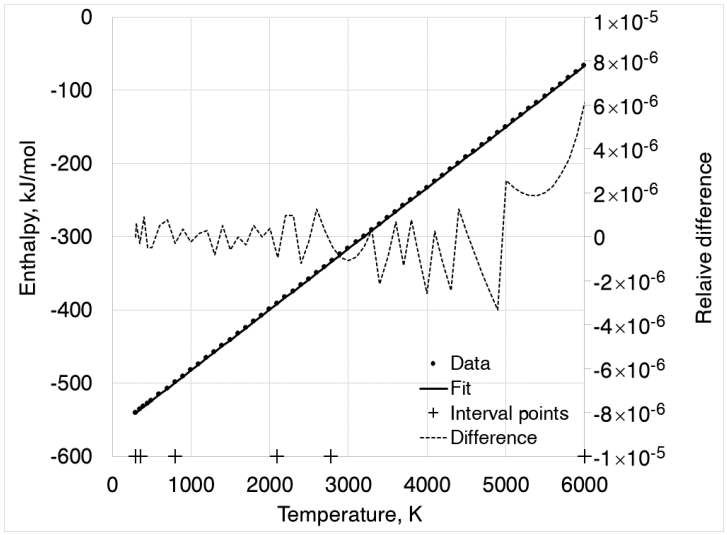


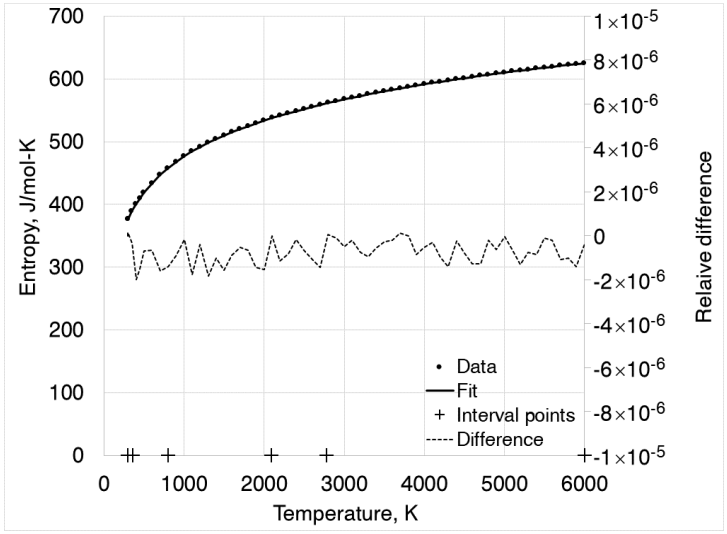


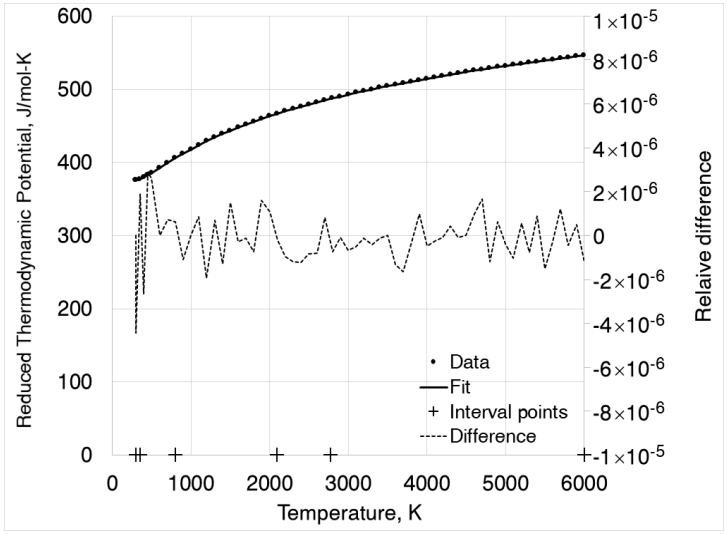
Potassium Bromide ($(KBr)_2$) – Br2K2(g)

MW = 238.005 g/mol – $\Delta_f H^o$(298.15 K) = -540.573 kJ/mol – S^o(298.15 K) = 376.23 J/mol-K – p^o = 0.1 MPa



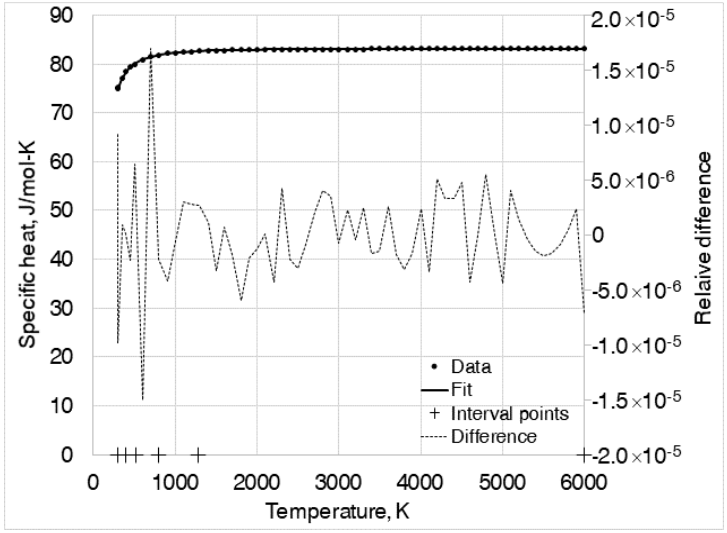


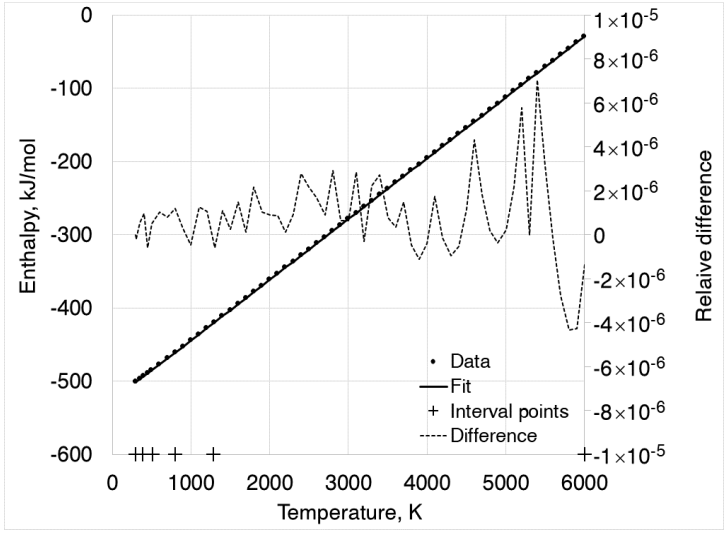


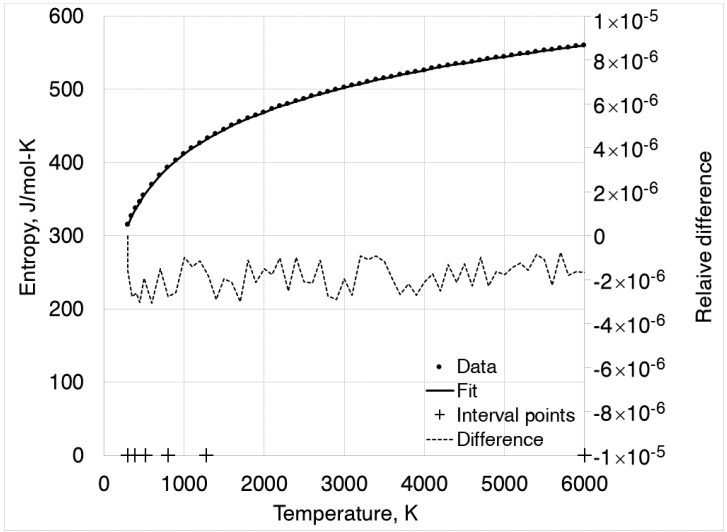


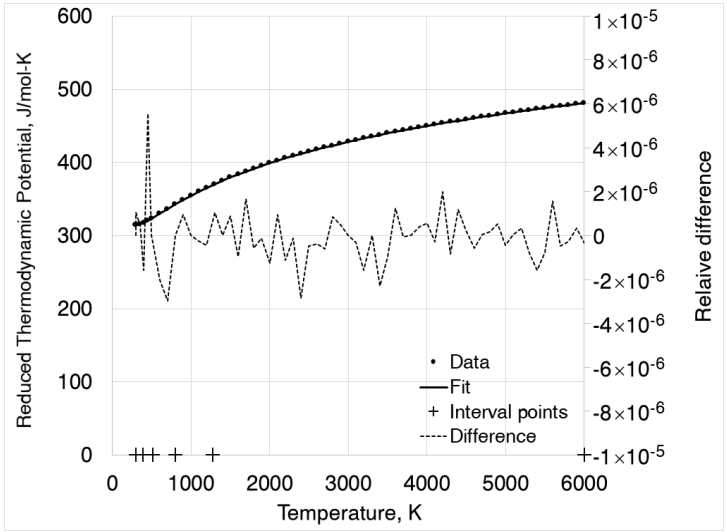
Lithium Bromide ($(LiBr)_2$) – Br2Li2(g)

MW = 173.69 g/mol – $\Delta_f H^o$(298.15 K) = -500.825 kJ/mol – S^o(298.15 K) = 314.534 J/mol-K – p^o = 0.1 MPa



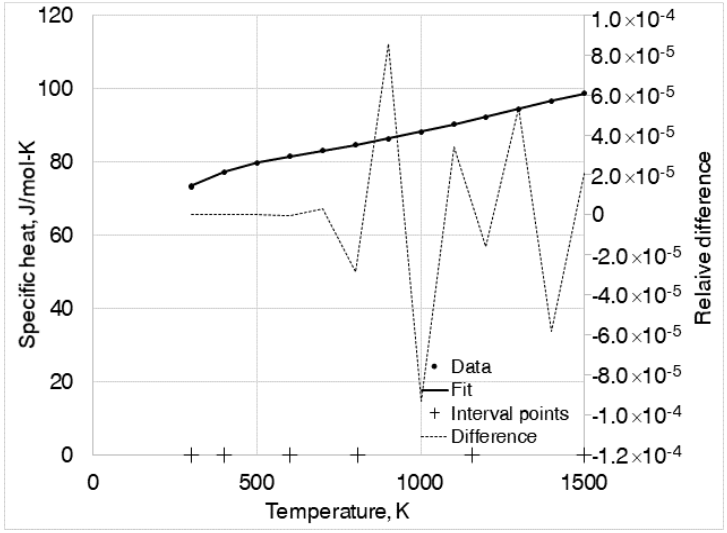


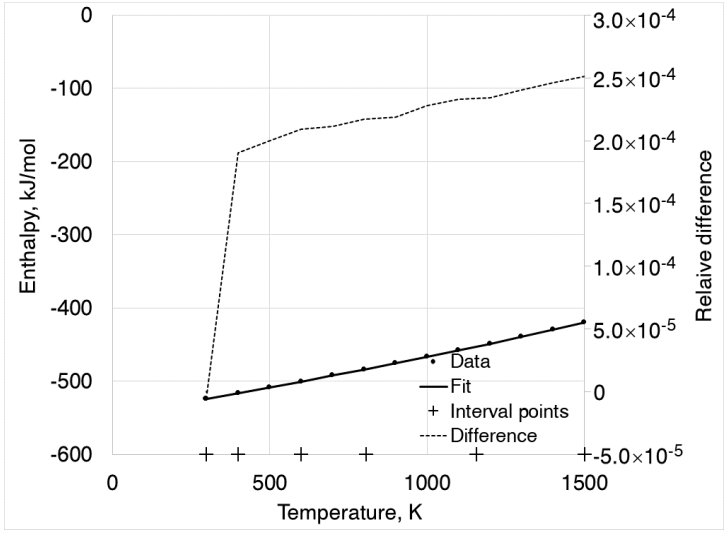


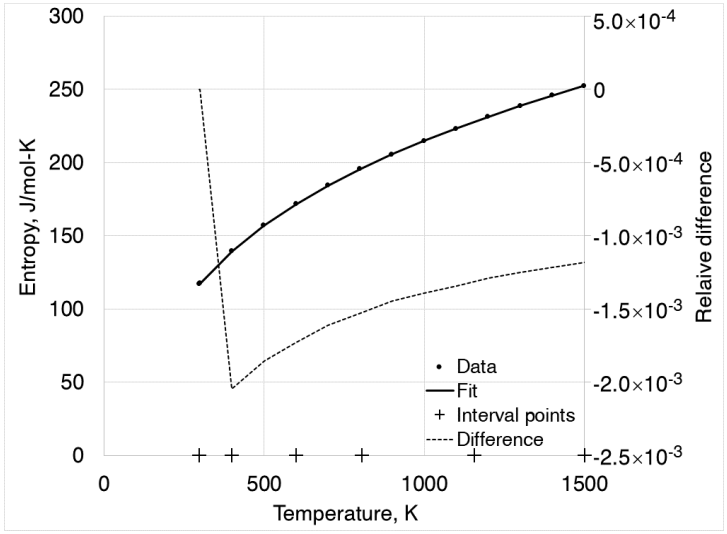


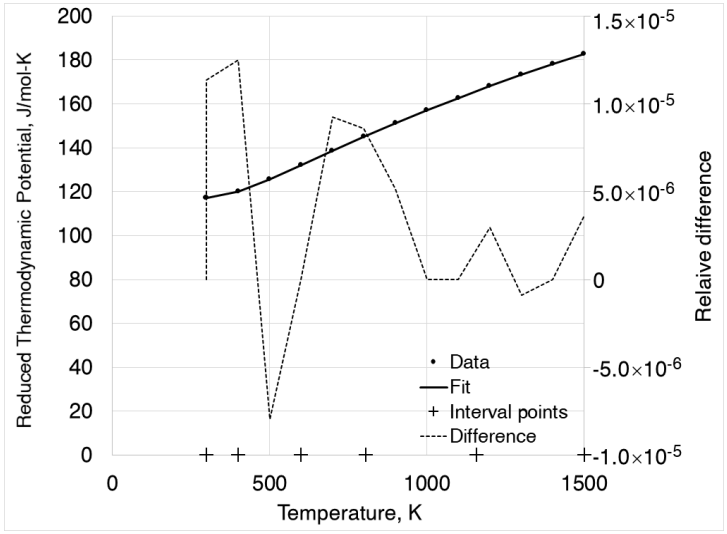
Magnesium Bromide ($MgBr_2$) – Br2Mg1(cr)

MW = 184.113 g/mol – $\Delta_f H^o$(298.15 K) = -524.255 kJ/mol – S^o(298.15 K) = 117.152 J/mol-K – p^o = 0.1 MPa



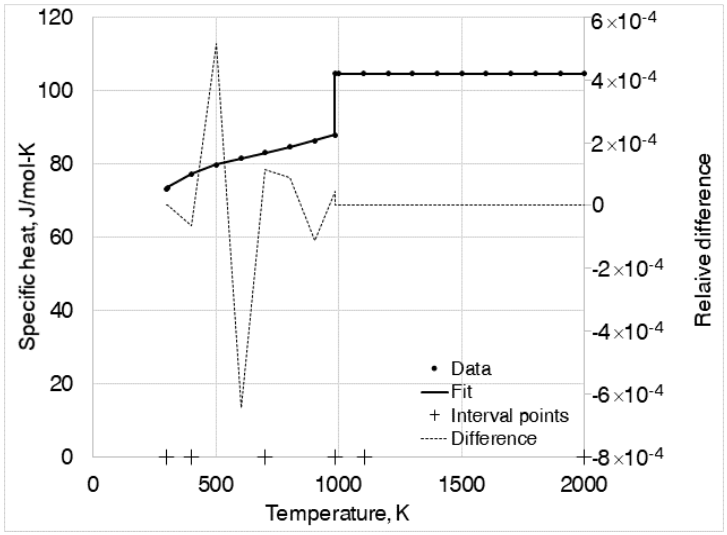


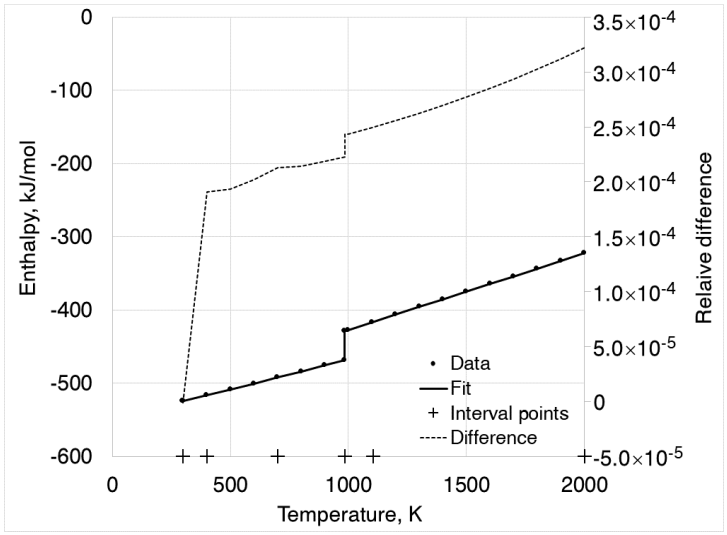


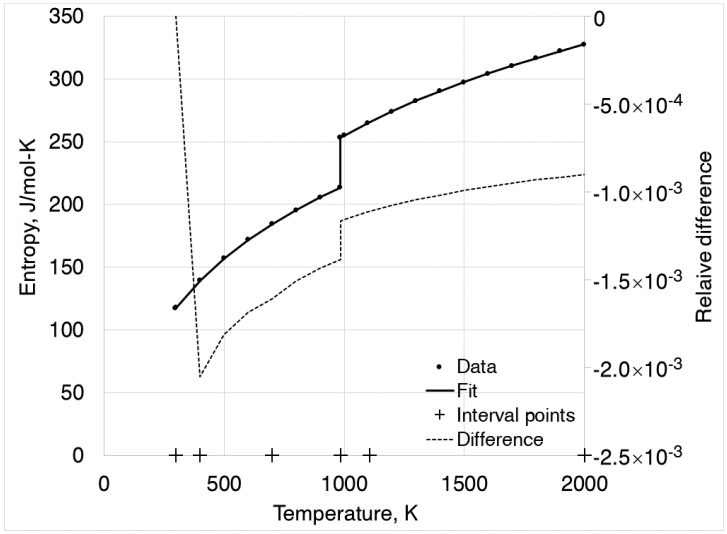


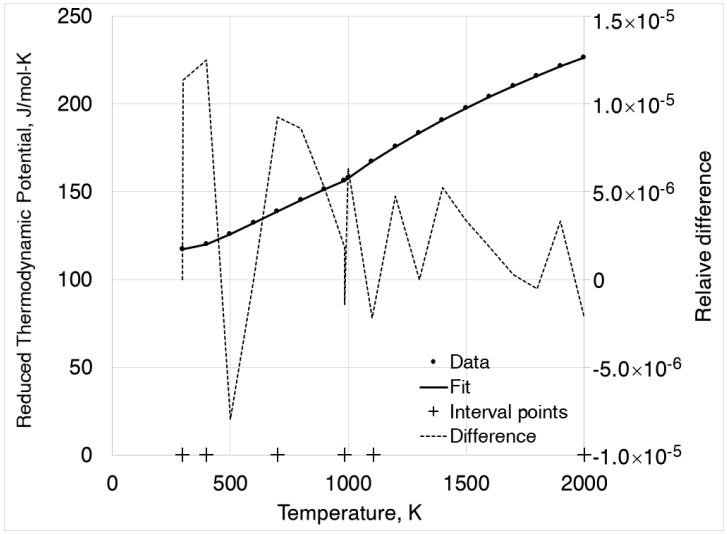
Magnesium Bromide ($MgBr_2$) – Br2Mg1(cr,l)

MW = 184.113 g/mol – $\Delta_f H^o$(298.15 K) = -524.255 kJ/mol – S^o(298.15 K) = 117.152 J/mol-K – p^o = 0.1 MPa



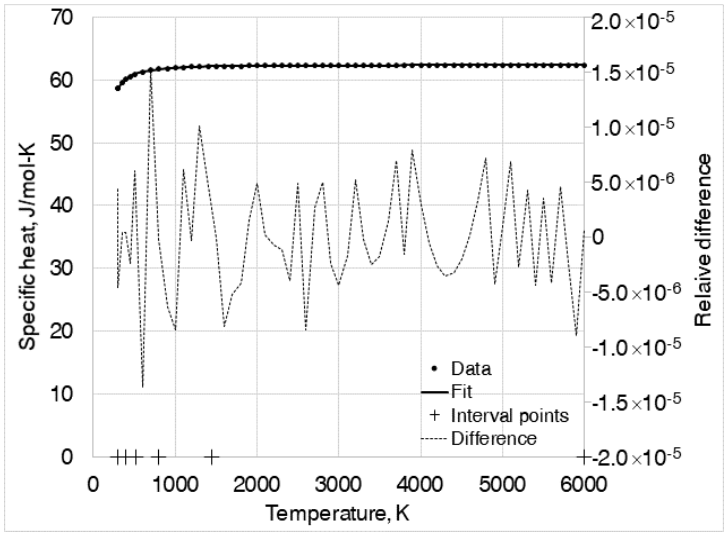


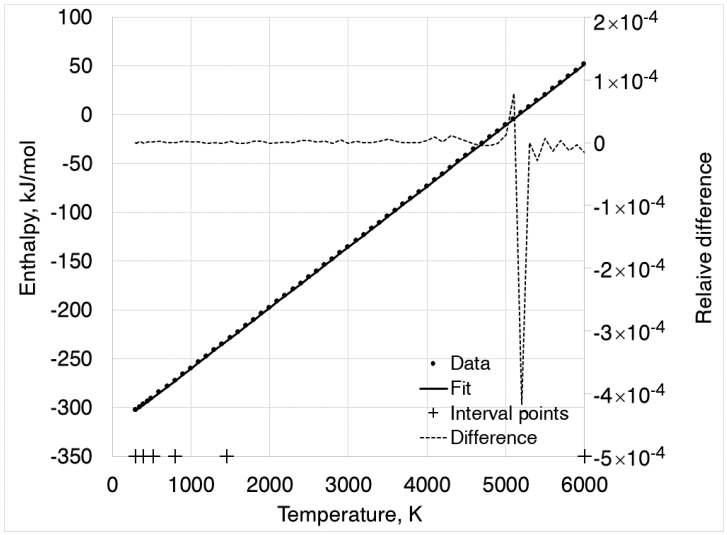


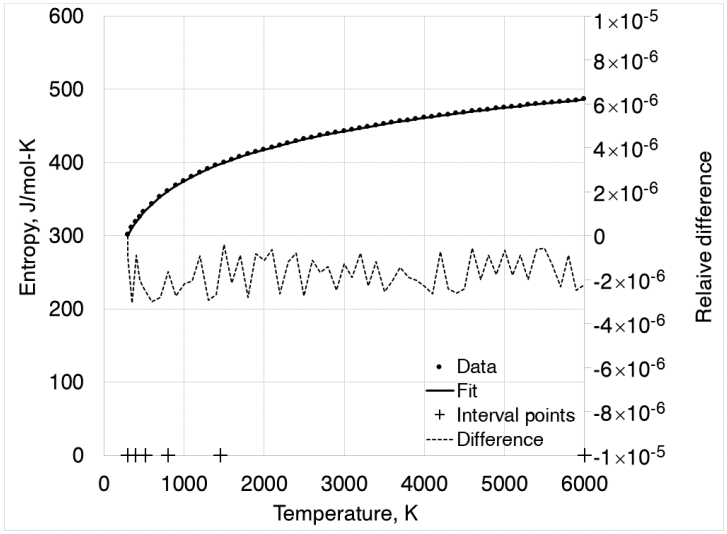


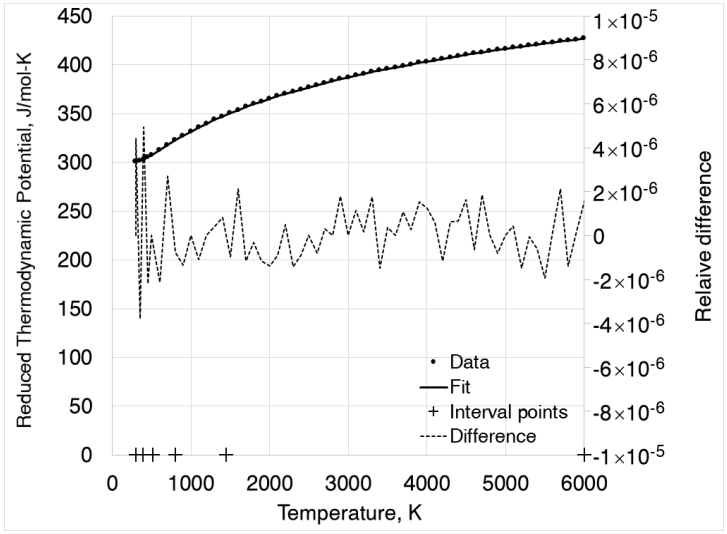
Magnesium Bromide ($MgBr_2$) – Br2Mg1(g)

MW = 184.113 g/mol – $\Delta_f H^o$(298.15 K) = -302.922 kJ/mol – S^o(298.15 K) = 301.038 J/mol-K – p^o = 0.1 MPa



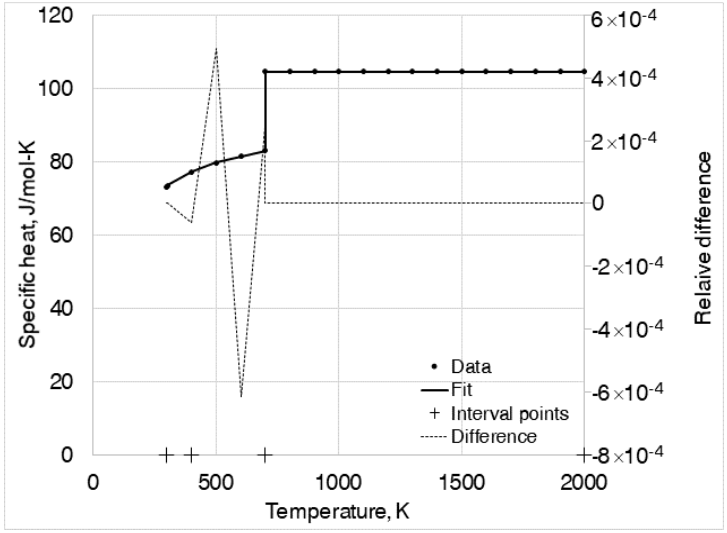


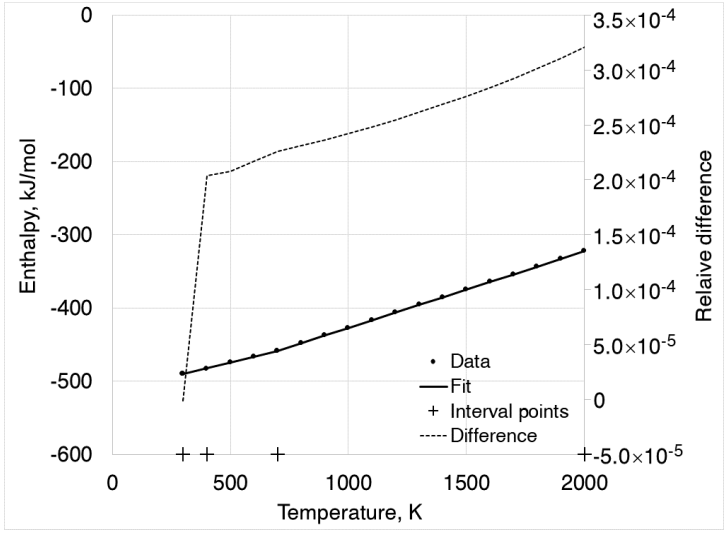


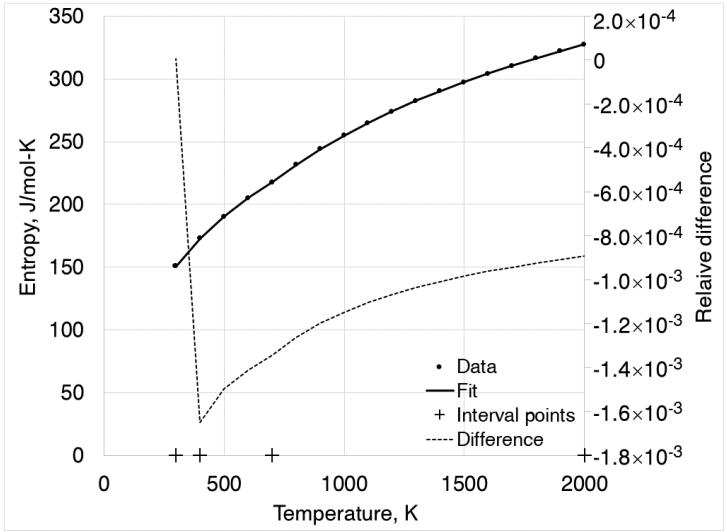


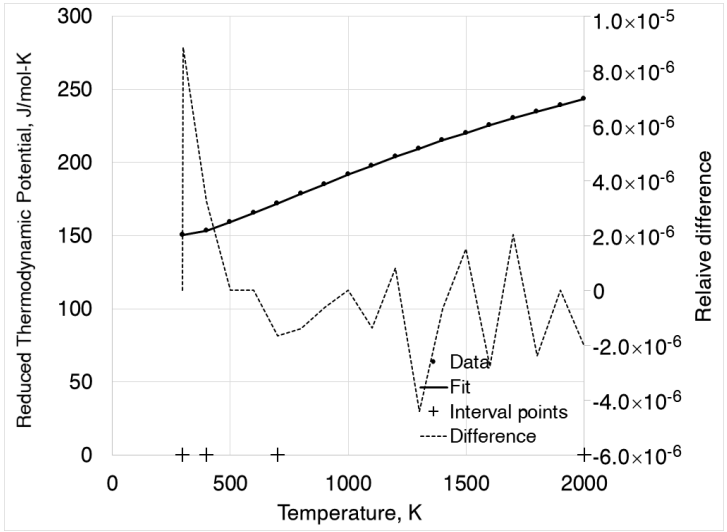
Magnesium Bromide ($MgBr_2$) – Br2Mg1(l)

MW = 184.113 g/mol – $\Delta_f H^o$(298.15 K) = -490.408 kJ/mol – S^o(298.15 K) = 150.5 J/mol-K – p^o = 0.1 MPa



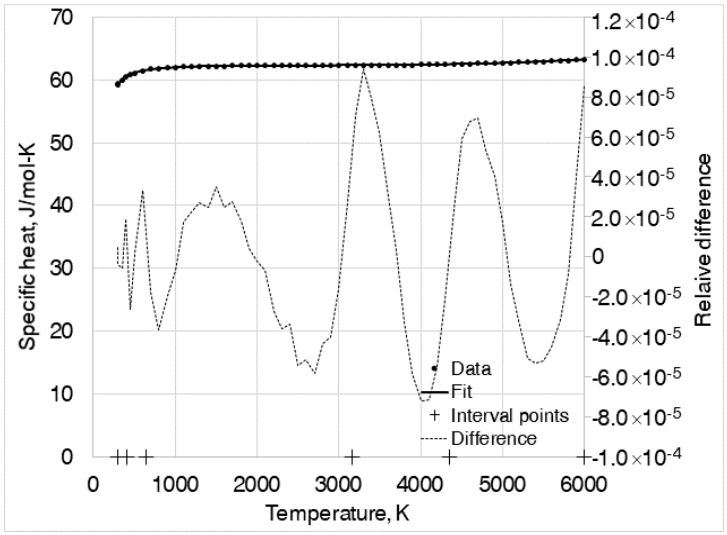


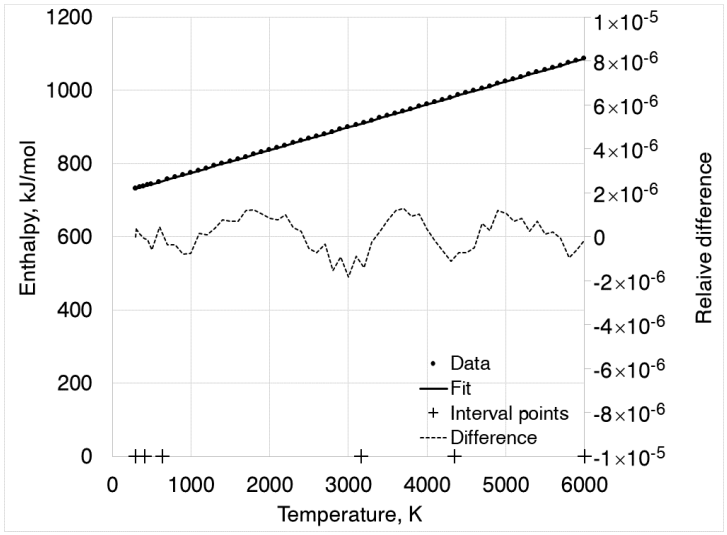


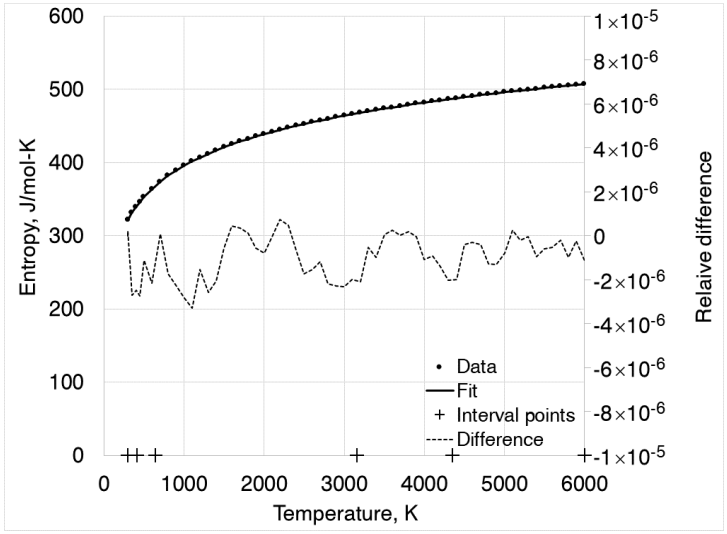


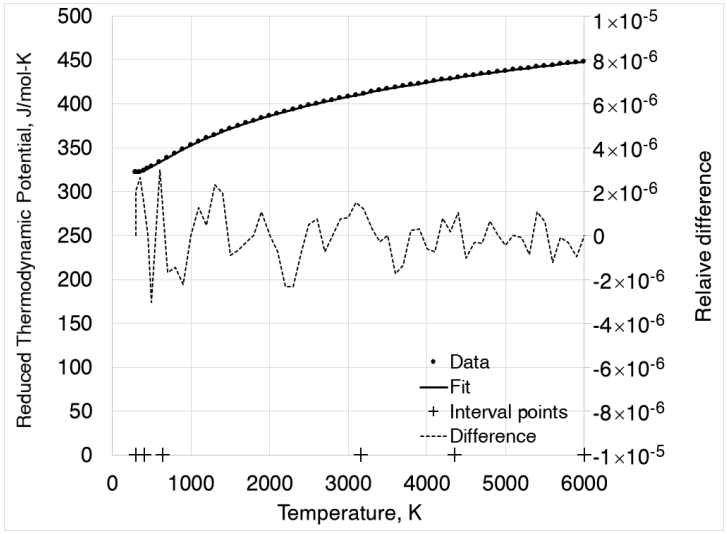
Magnesium Bromide, Ion ($MgBr_2^+$) – Br2Mg1+(g)

MW = 184.112 g/mol – $\Delta_f H^o$(298.15 K) = 731.363 kJ/mol – S^o(298.15 K) = 321.737 J/mol-K – p^o = 0.1 MPa



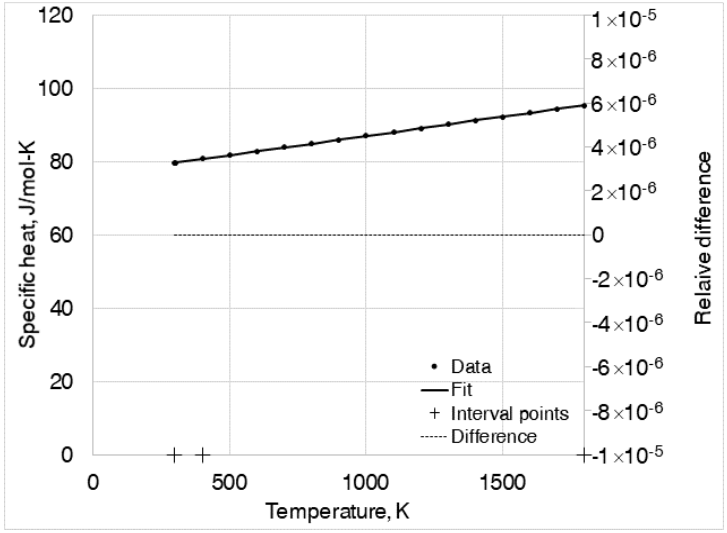


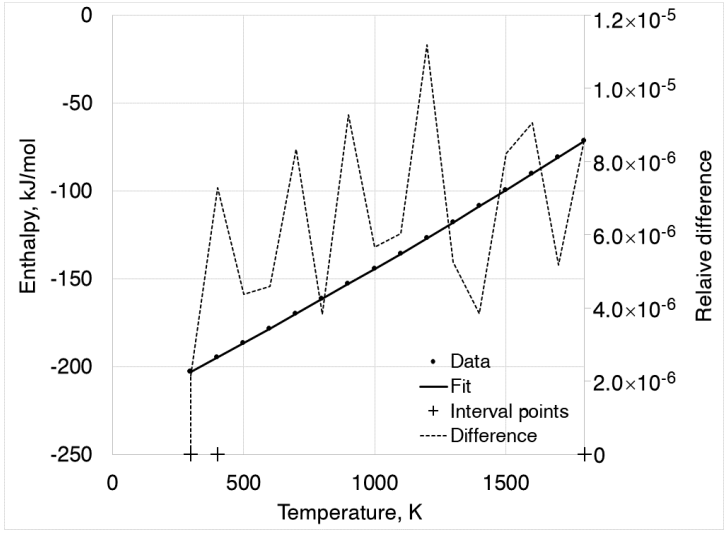


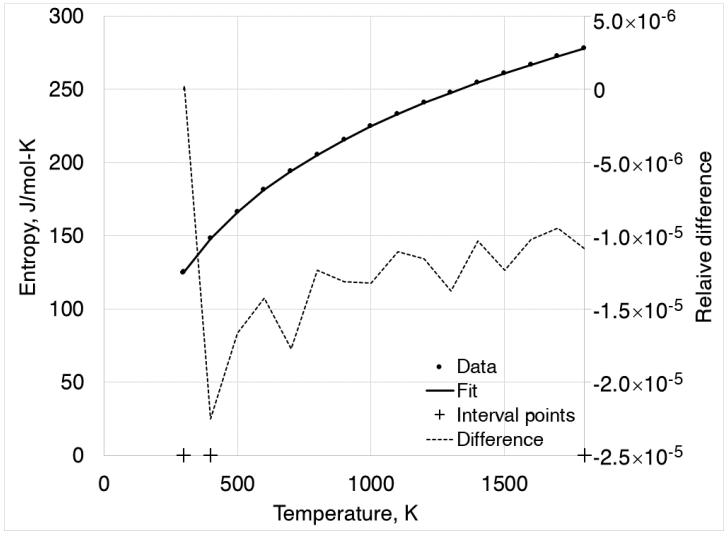


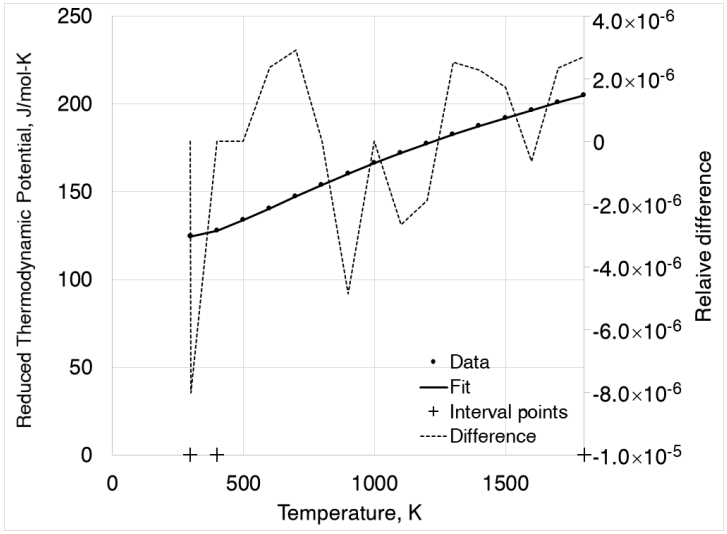
Molybdenum Bromide ($MoBr_2$) – Br2Mo1(cr)

MW = 255.75 g/mol – $\Delta_f H^o$(298.15 K) = -202.924 kJ/mol – S^o(298.15 K) = 124.683 J/mol-K – p^o = 0.1 MPa



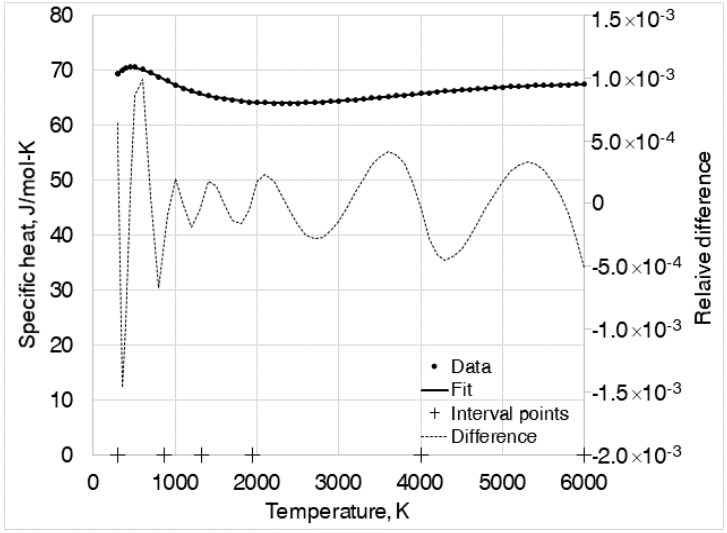


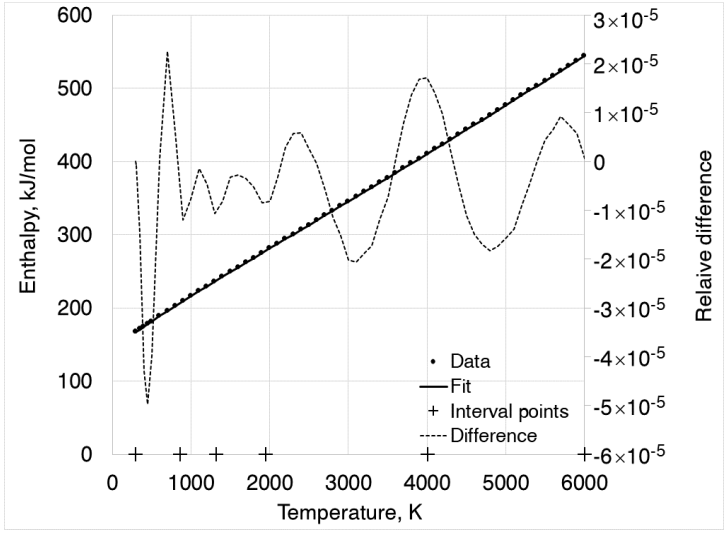


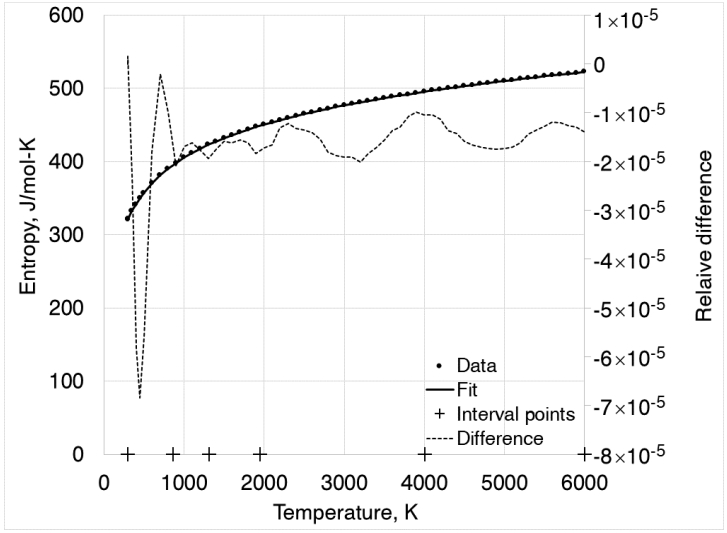


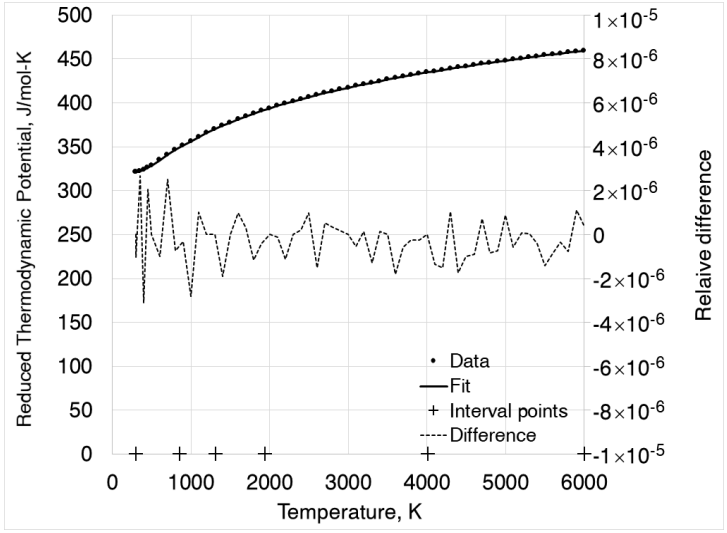
Molybdenum Bromide ($MoBr_2$) – Br2Mo1(g)

MW = 255.75 g/mol – $\Delta_f H^o$(298.15 K) = 167.36 kJ/mol – S^o(298.15 K) = 320.96 J/mol-K – p^o = 0.1 MPa



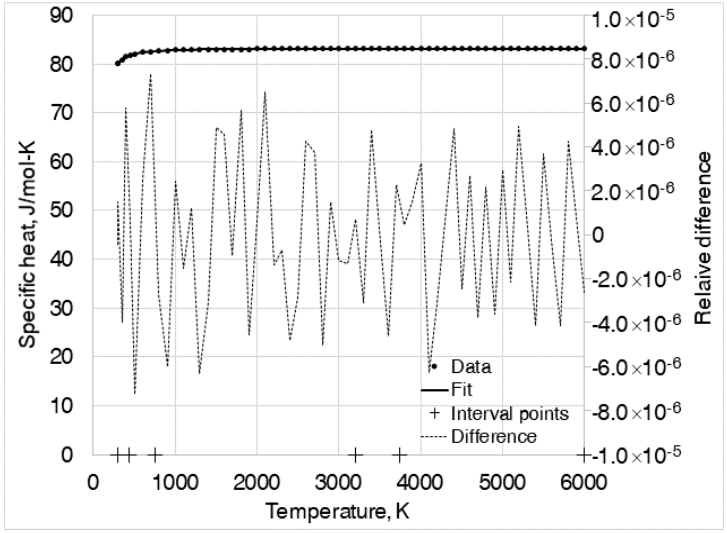


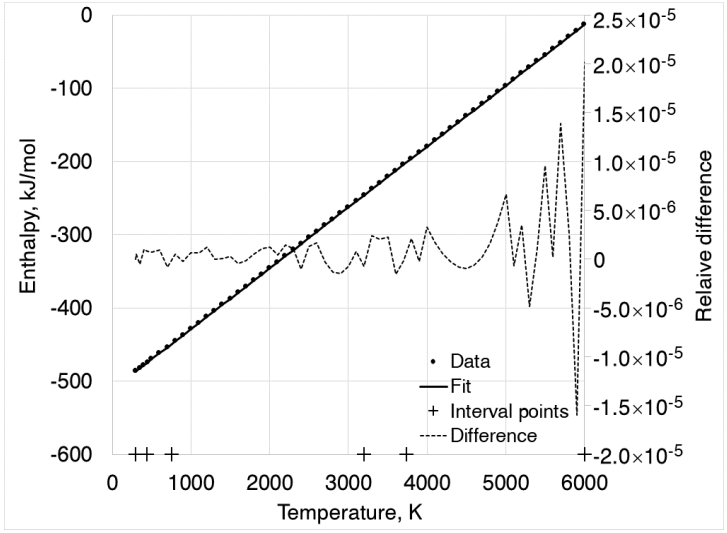


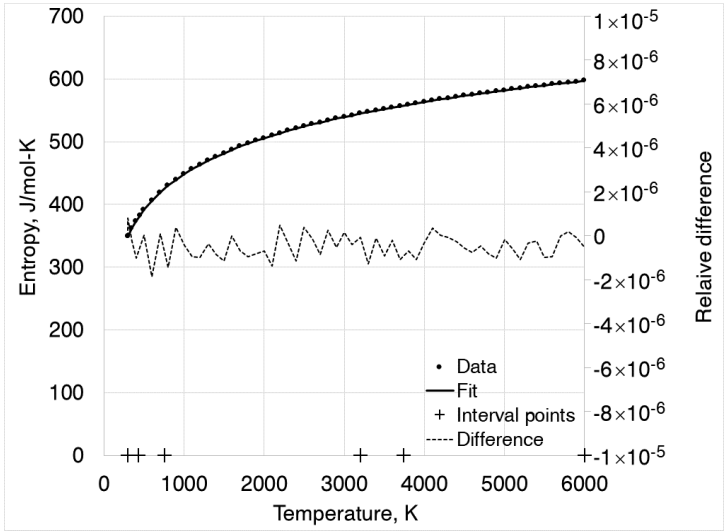


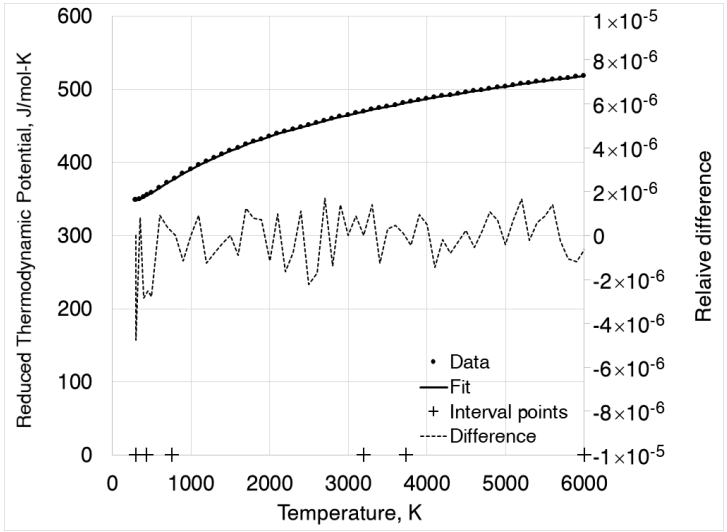
Sodium Bromide ((NaBr)$_2$) – Br2Na2(g)

MW = 205.788 g/mol – $\Delta_f H^o$(298.15 K) = -486.348 kJ/mol – S^o(298.15 K) = 348.949 J/mol-K – p^o = 0.1 MPa



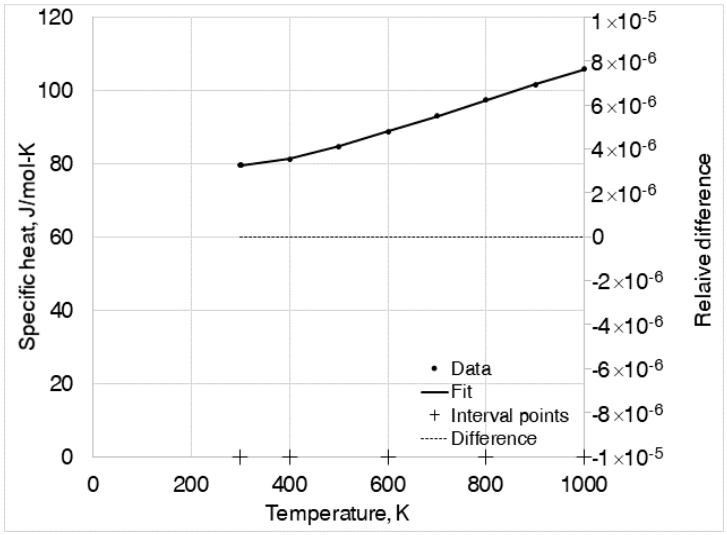


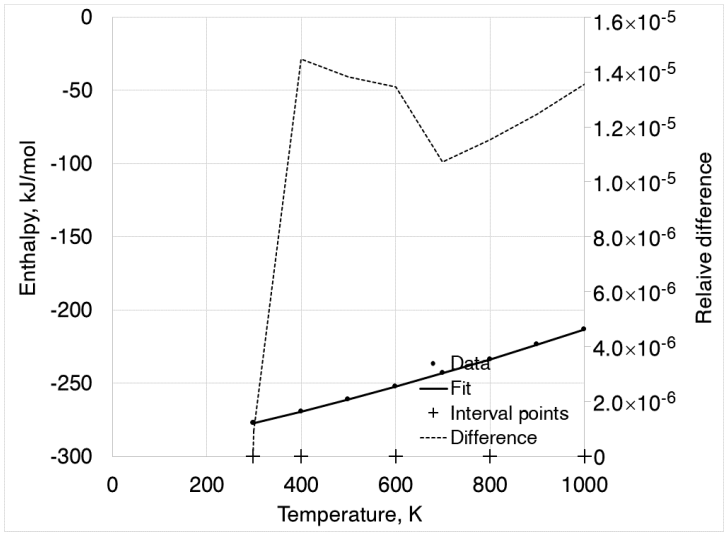


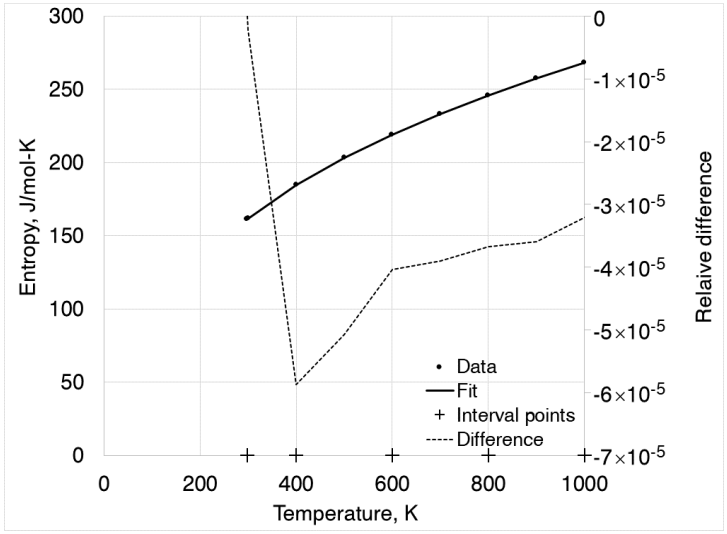


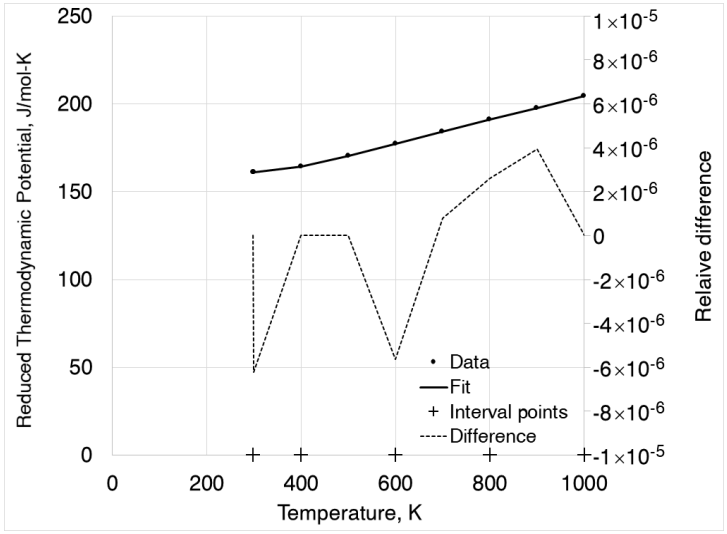
Lead Bromide ($PbBr_2$) – Br2Pb1(cr)

MW = 367 g/mol – $\Delta_f H^o$(298.15 K) = -277.399 kJ/mol – S^o(298.15 K) = 161.128 J/mol-K – p^o = 0.1 MPa



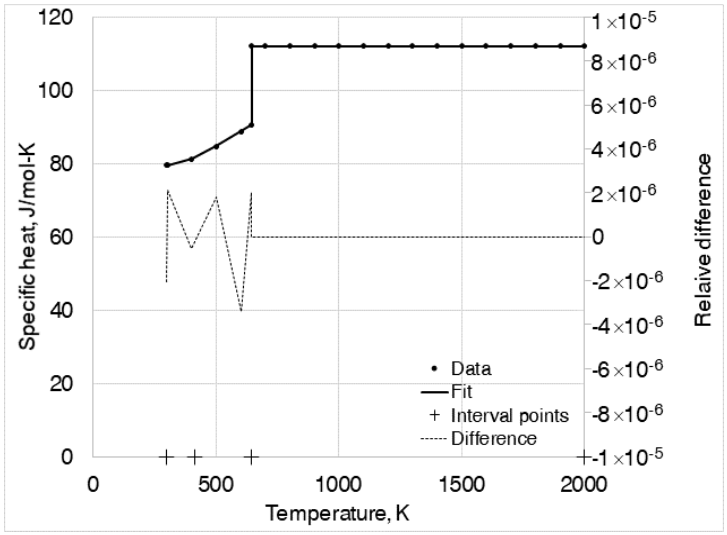


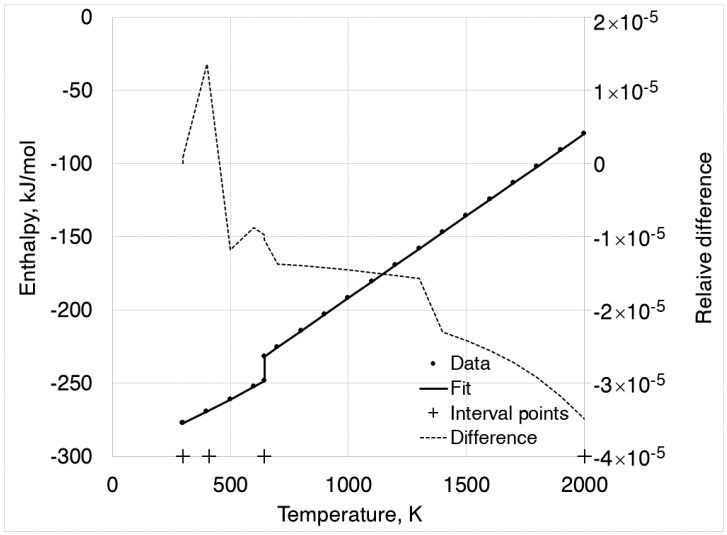


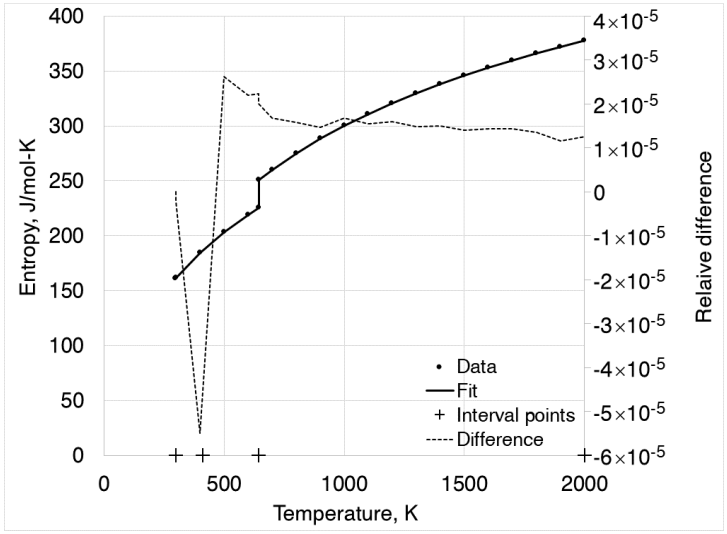


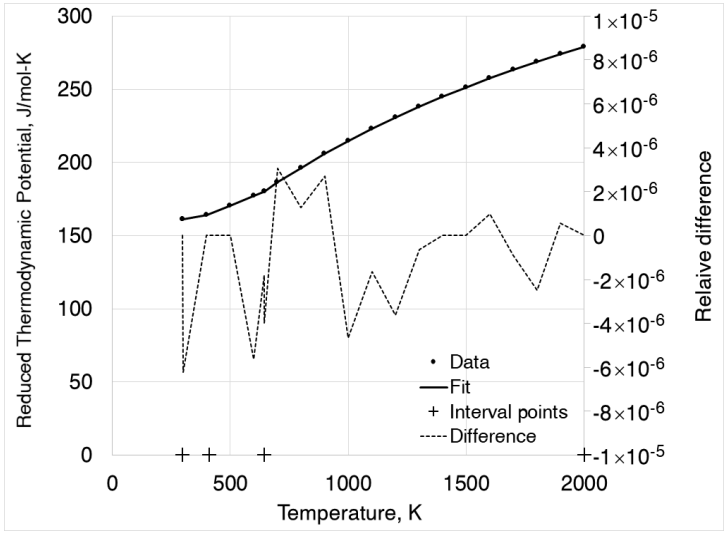
Lead Bromide ($PbBr_2$) – Br2Pb1(cr,l)

MW = 367 g/mol – $\Delta_f H^o$(298.15 K) = -277.399 kJ/mol – S^o(298.15 K) = 161.128 J/mol-K – p^o = 0.1 MPa



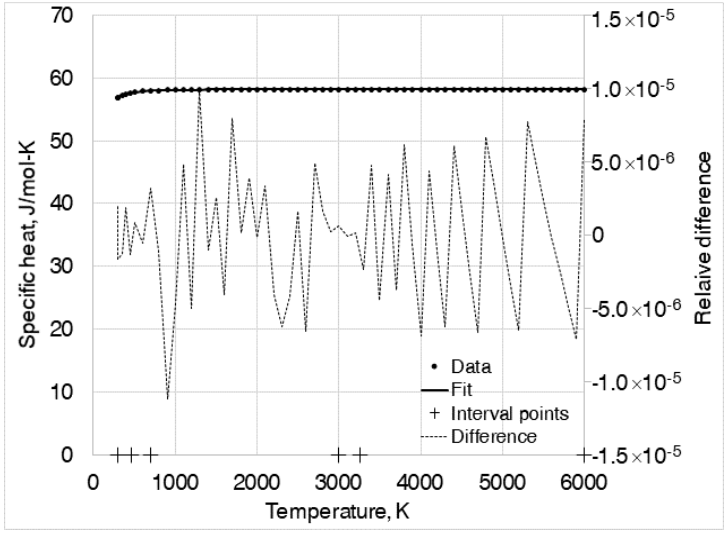


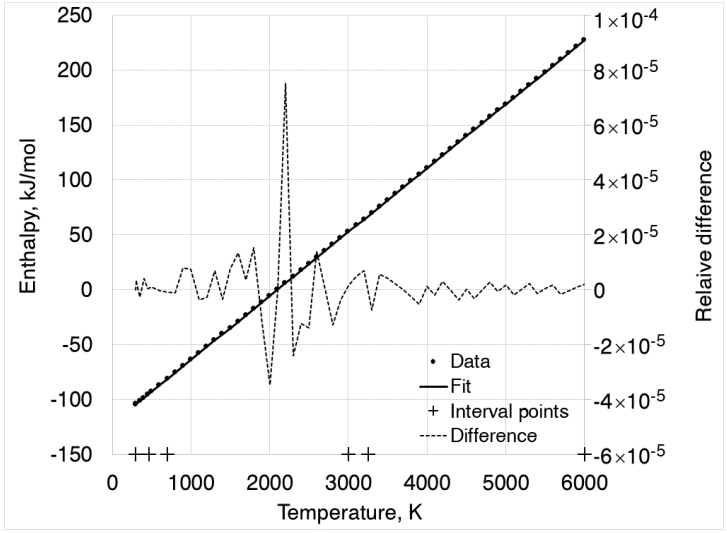


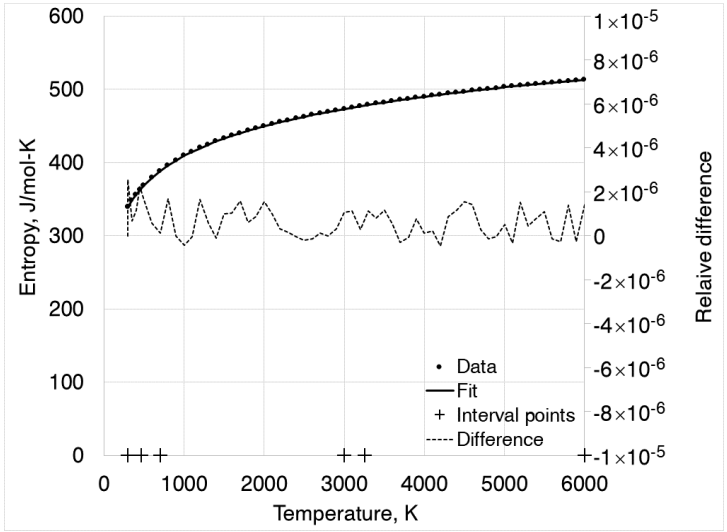


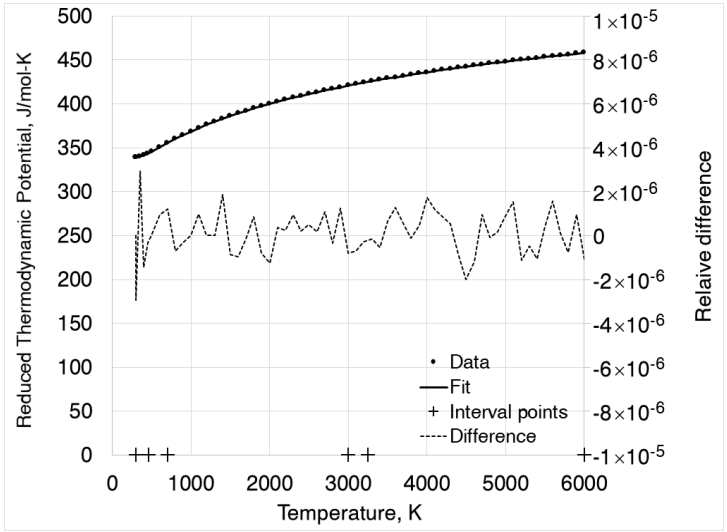
Lead Bromide ($PbBr_2$) – Br2Pb1(g)

MW = 367 g/mol – $\Delta_f H^o$(298.15 K) = -104.391 kJ/mol – S^o(298.15 K) = 339.386 J/mol-K – p^o = 0.1 MPa



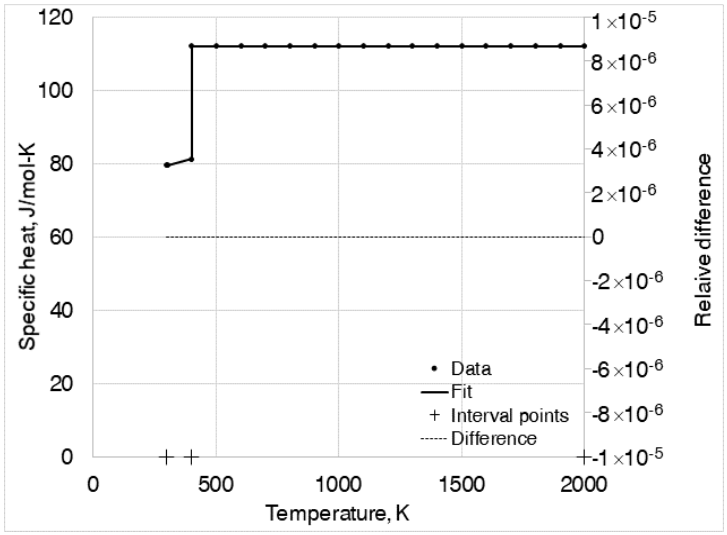


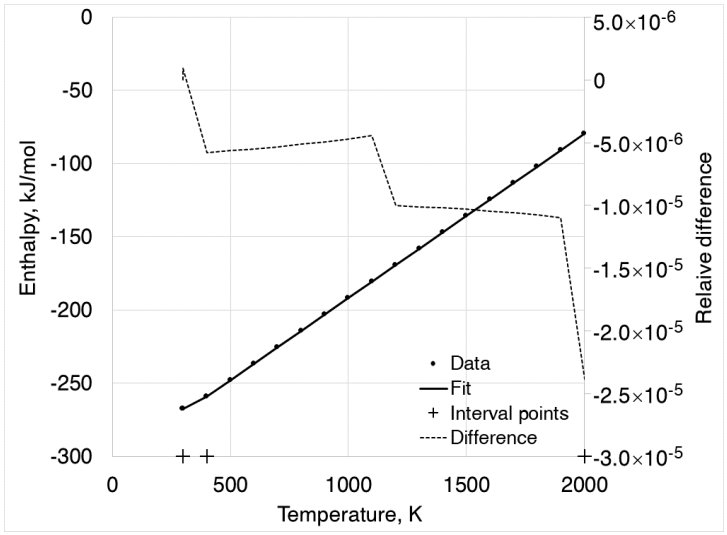


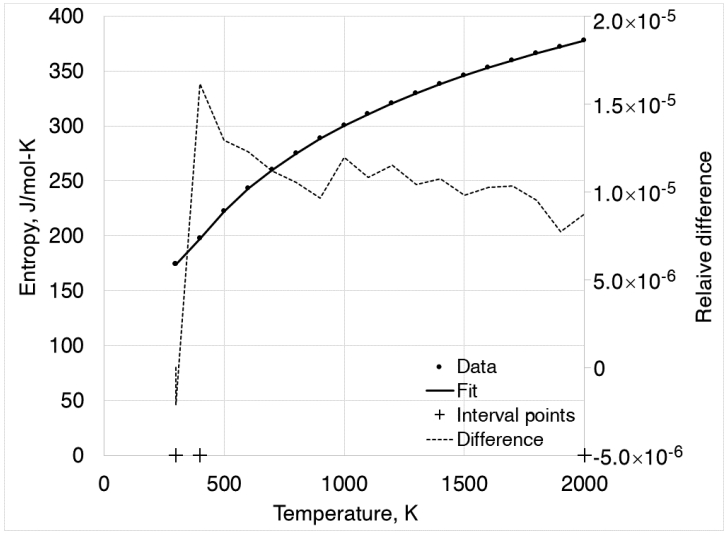


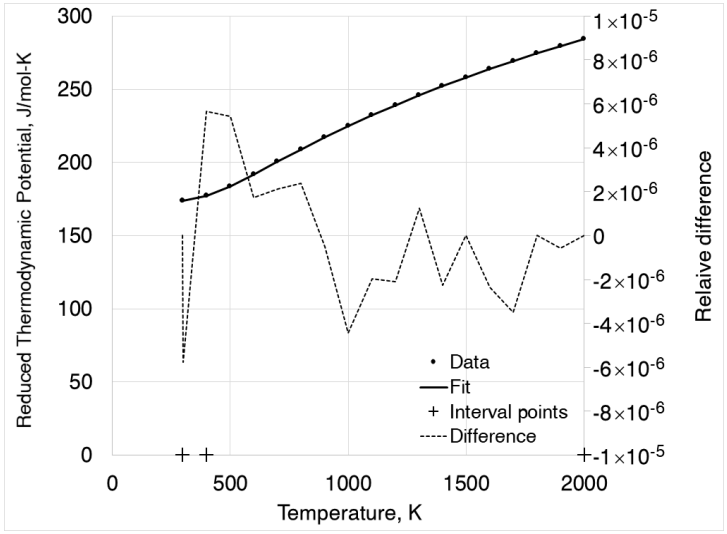
Lead Bromide ($PbBr_2$) – Br2Pb1(l)

MW = 367 g/mol – $\Delta_f H^o$(298.15 K) = -267.412 kJ/mol – S^o(298.15 K) = 173.878 J/mol-K – p^o = 0.1 MPa



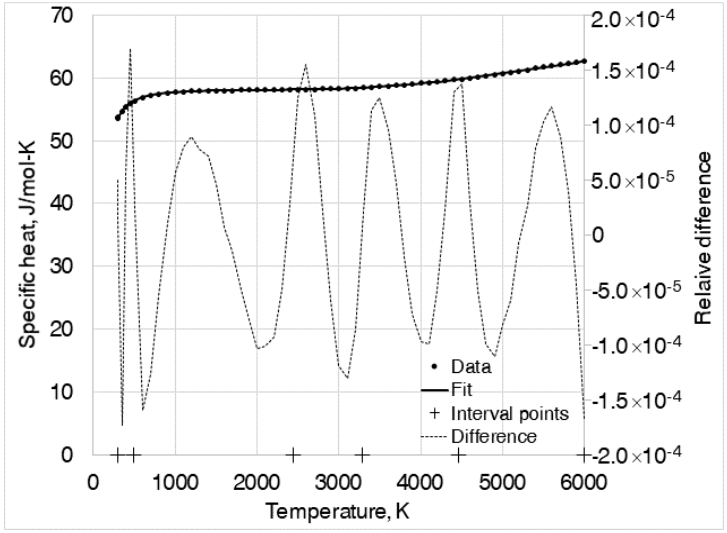


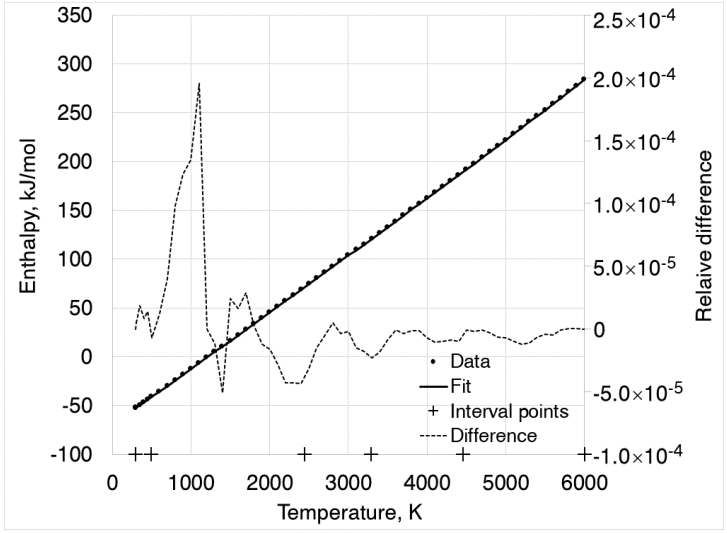


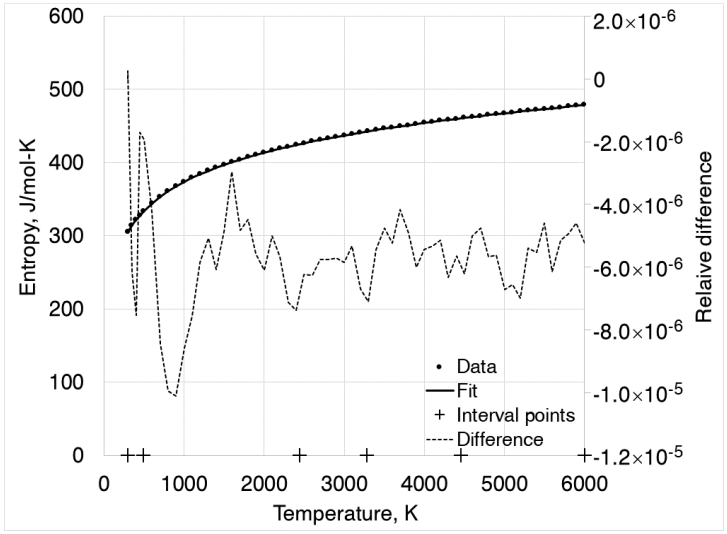


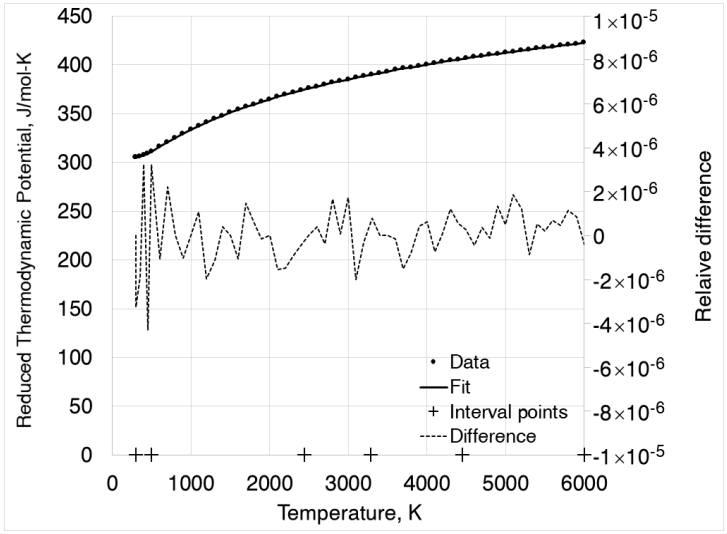
Dibromosilylene ($SiBr_2$) – Br2Si1(g)

MW = 187.894 g/mol – $\Delta_f H^o$(298.15 K) = -52.3 kJ/mol – S^o(298.15 K) = 305.225 J/mol-K – p^o = 0.1 MPa



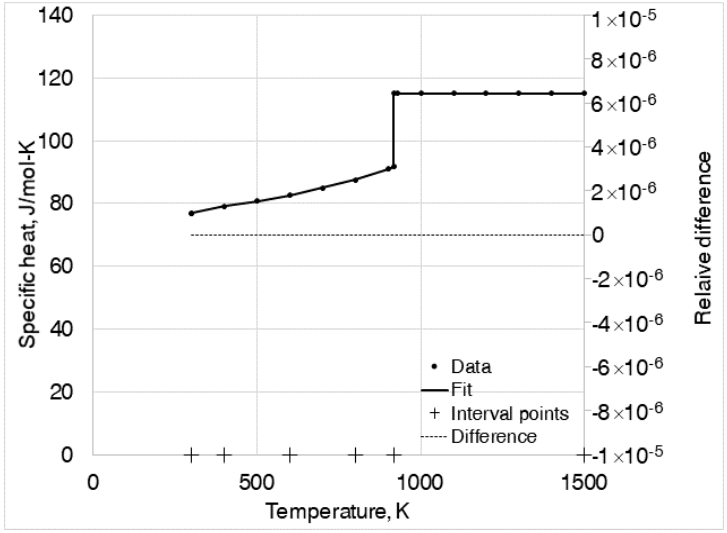


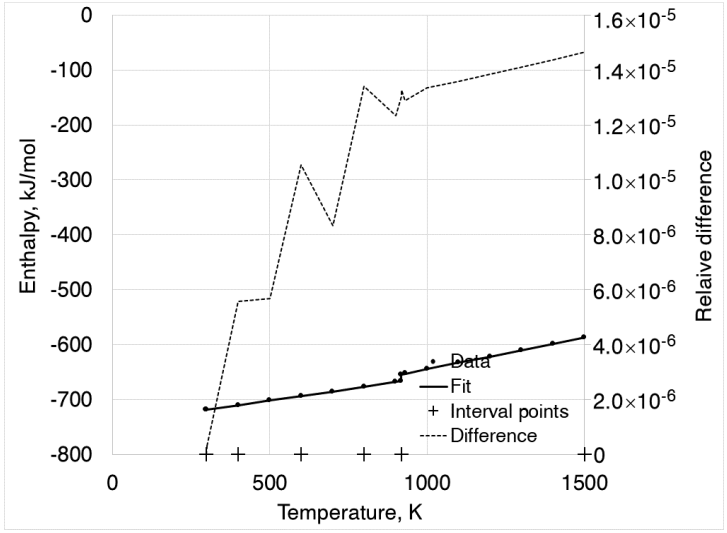


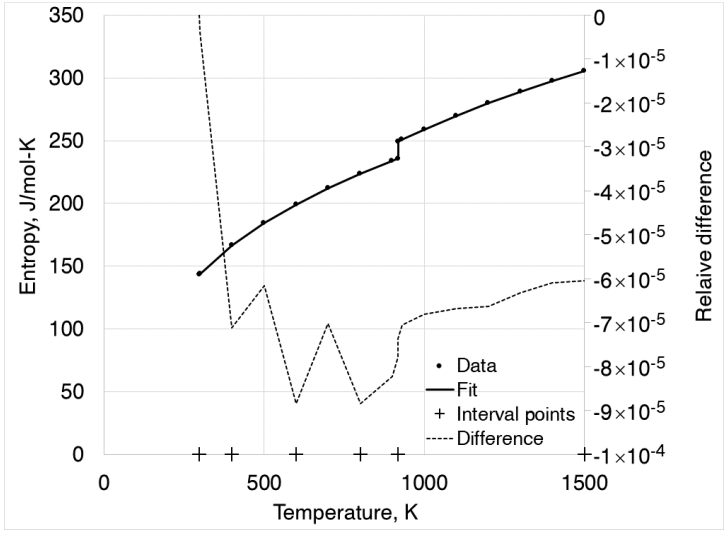


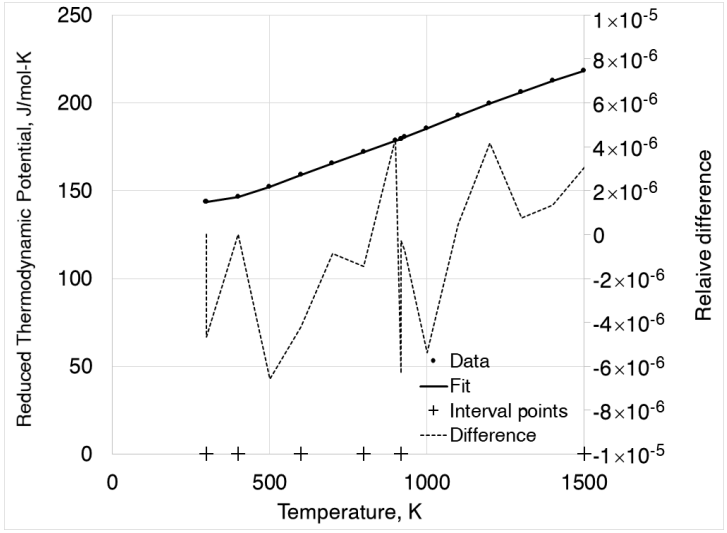
Strontium Bromide ($SrBr_2$) – Br2Sr1(cr)

MW = 247.43 g/mol – $\Delta_f H^o$(298.15 K) = -717.974 kJ/mol – S^o(298.15 K) = 143.438 J/mol-K – p^o = 0.1 MPa



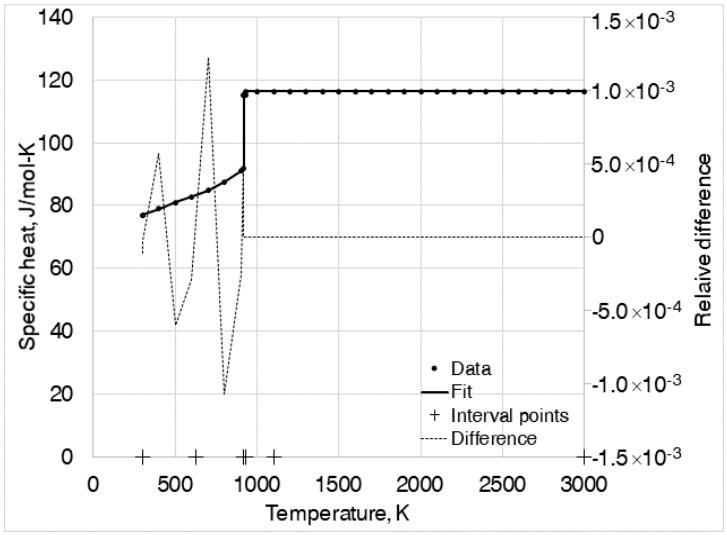


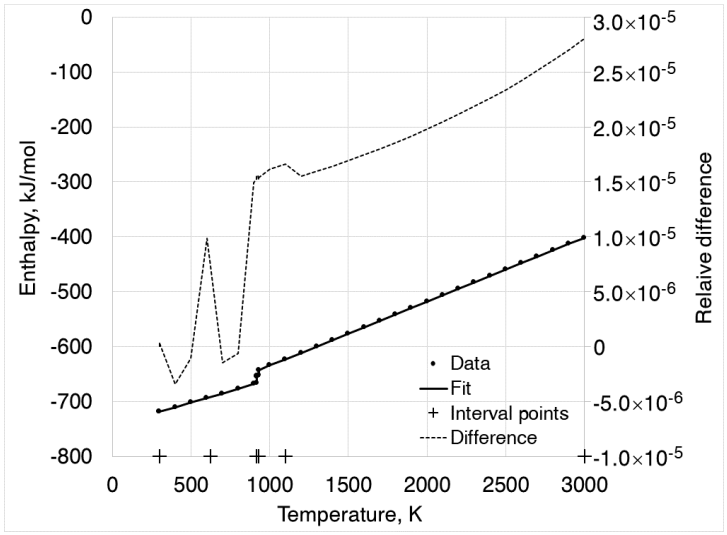


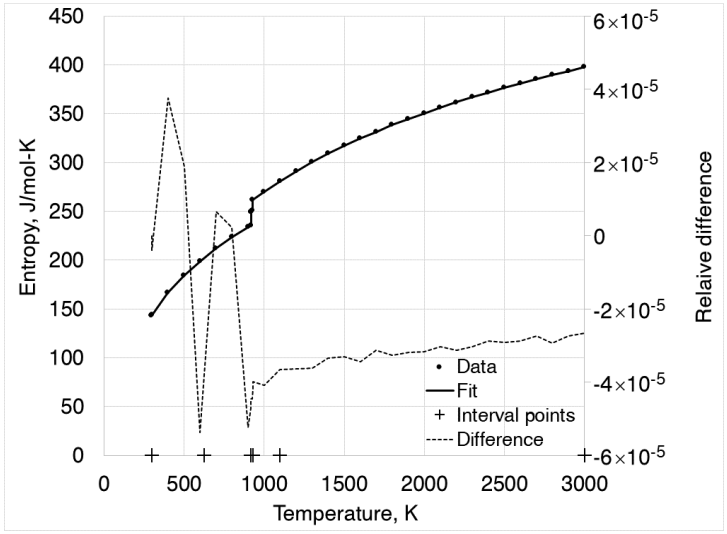


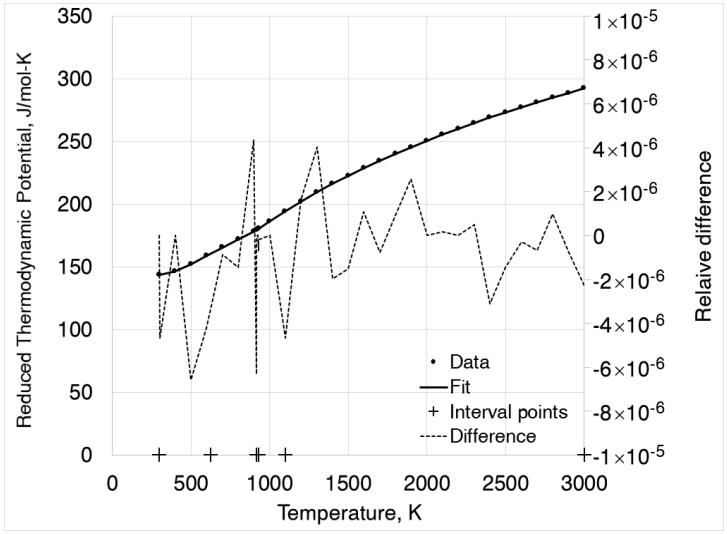
Strontium Bromide ($SrBr_2$) – Br2Sr1(cr,l)

MW = 247.43 g/mol – $\Delta_f H^o$(298.15 K) = -717.974 kJ/mol – S^o(298.15 K) = 143.438 J/mol-K – p^o = 0.1 MPa



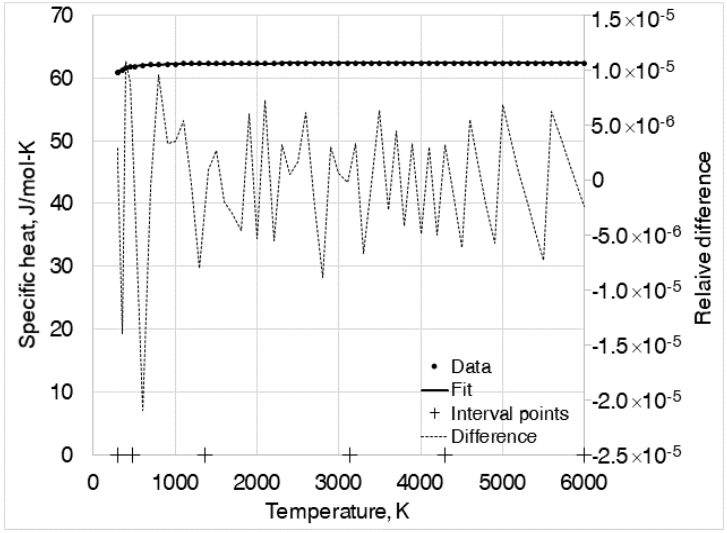


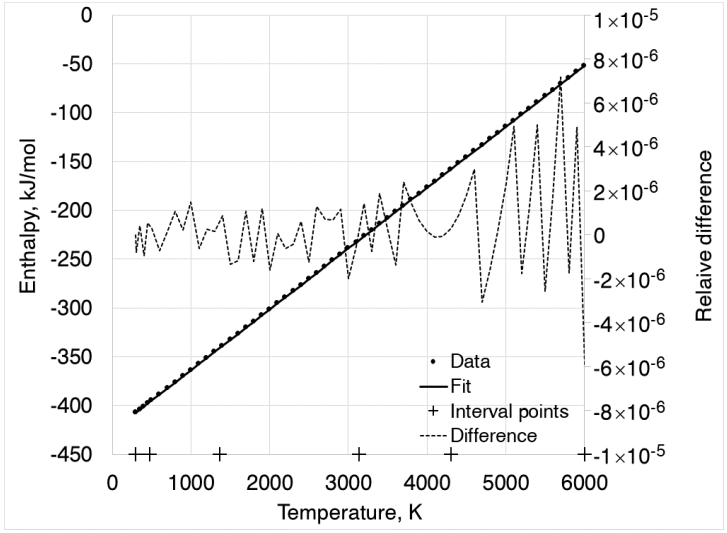


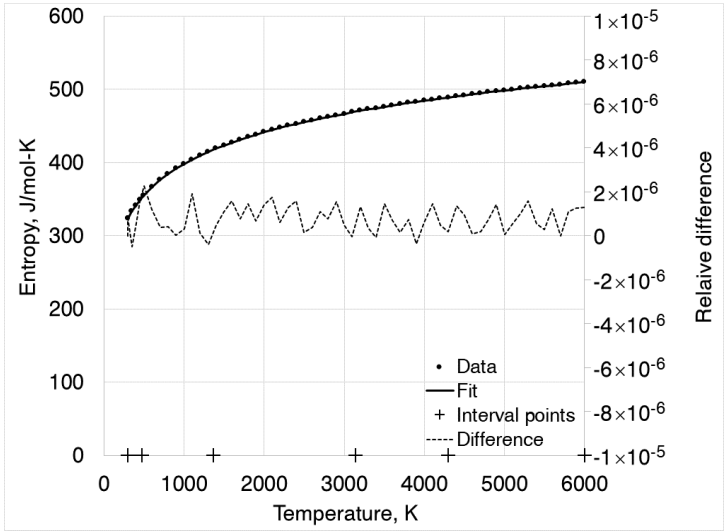


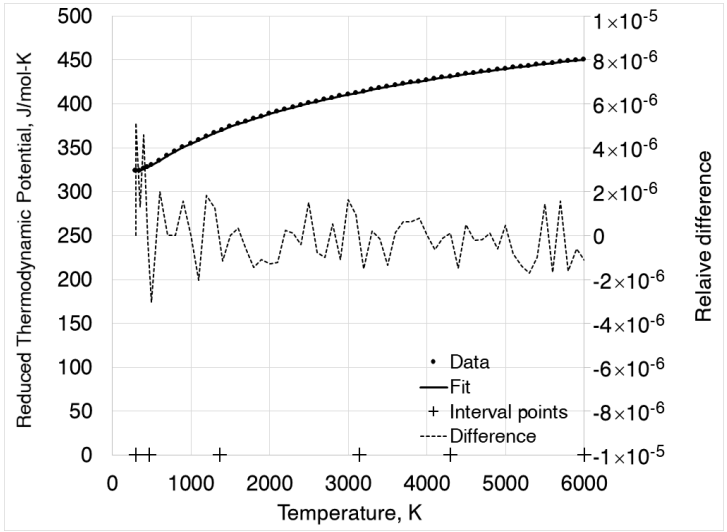
Strontium Bromide ($SrBr_2$) – Br2Sr1(g)

MW = 247.43 g/mol – $\Delta_f H^o$(298.15 K) = -407.103 kJ/mol – S^o(298.15 K) = 323.413 J/mol-K – p^o = 0.1 MPa



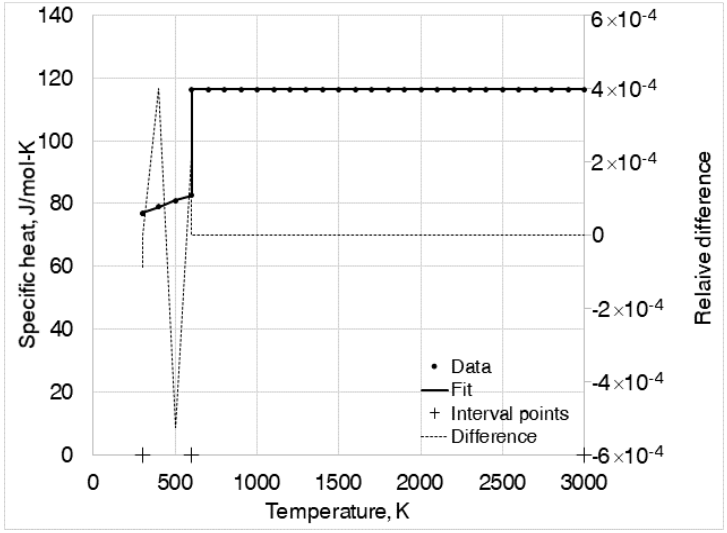


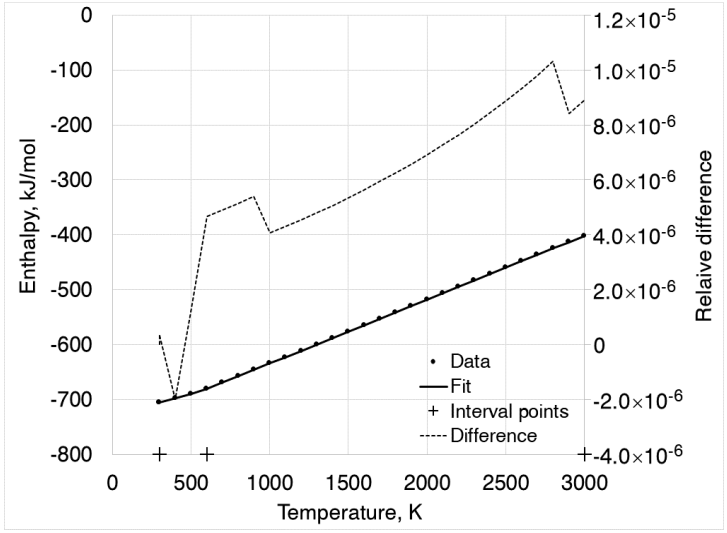


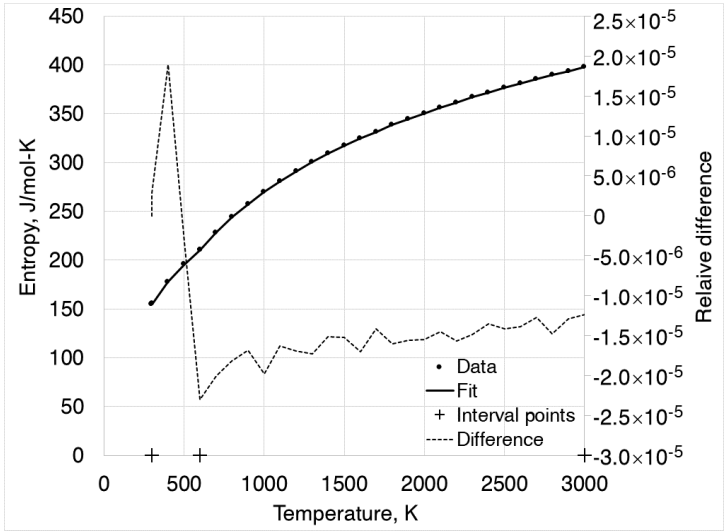


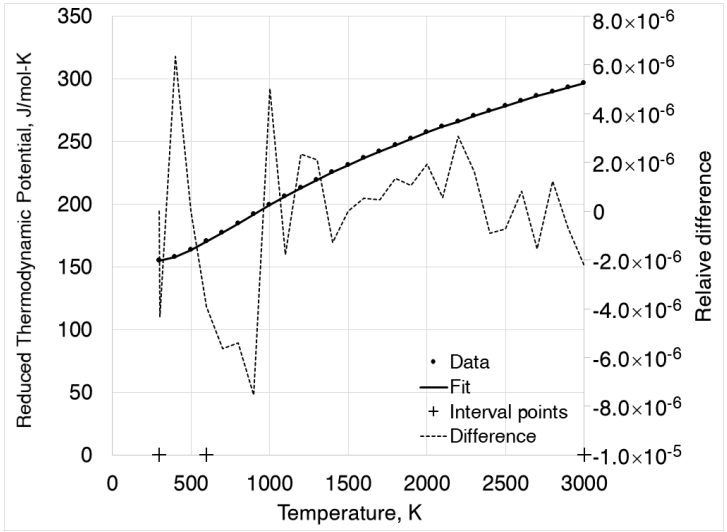
Strontium Bromide ($SrBr_2$) – Br2Sr1(l)

MW = 247.43 g/mol – $\Delta_f H^o$(298.15 K) = -705.098 kJ/mol – S^o(298.15 K) = 154.843 J/mol-K – p^o = 0.1 MPa



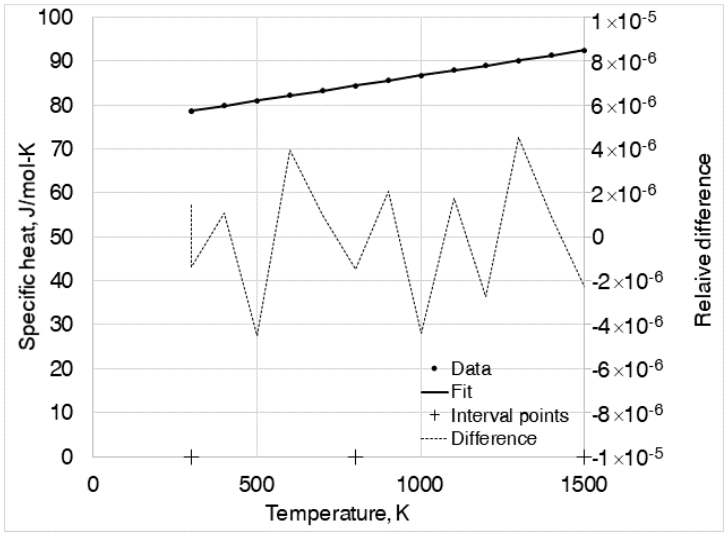


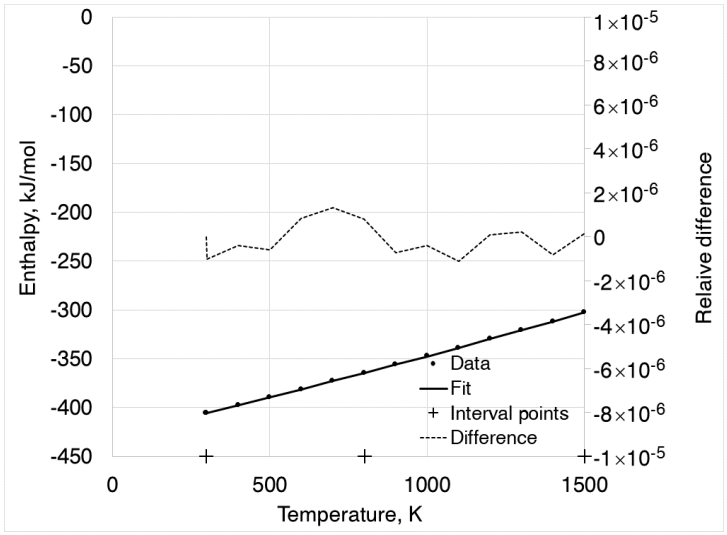


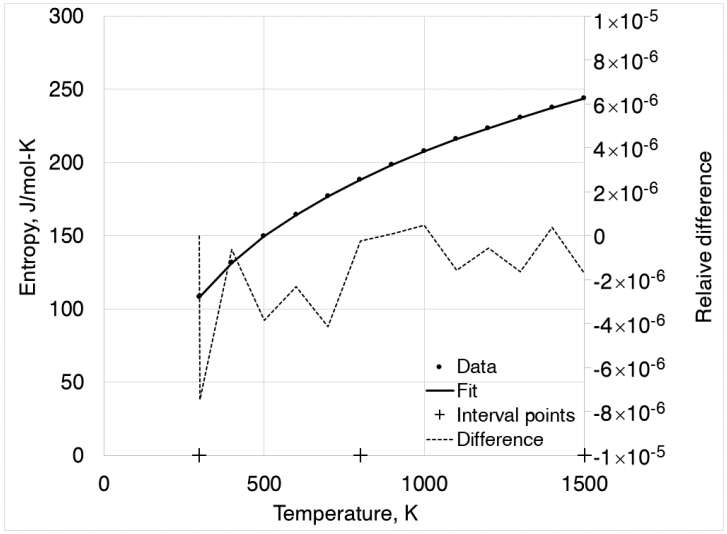


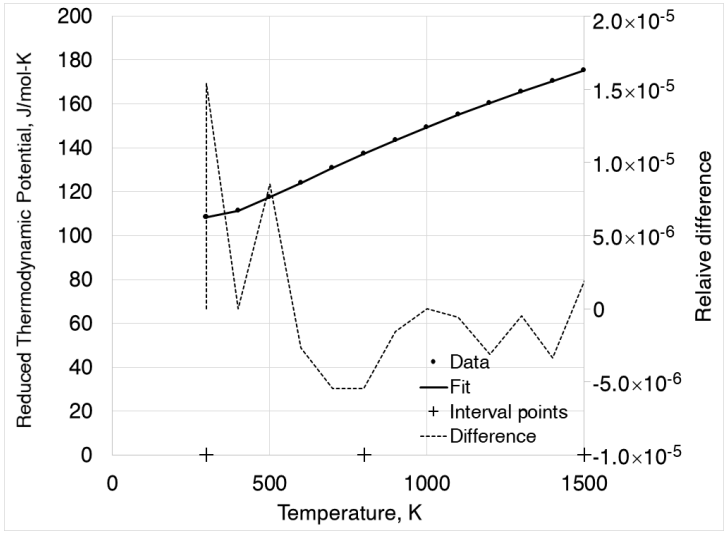
Titanium Bromide ($TiBr_2$) – Br2Ti1(cr)

MW = 207.69 g/mol – $\Delta_f H^o$(298.15 K) = -405.43 kJ/mol – S^o(298.15 K) = 108.366 J/mol-K – p^o = 0.1 MPa



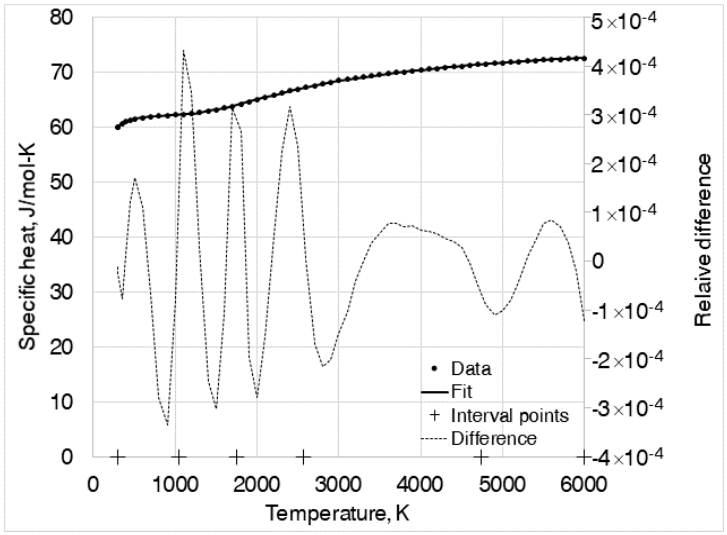


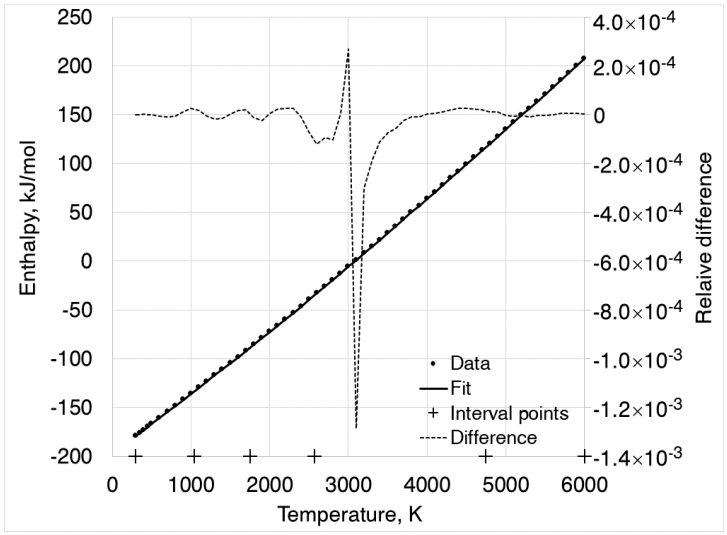


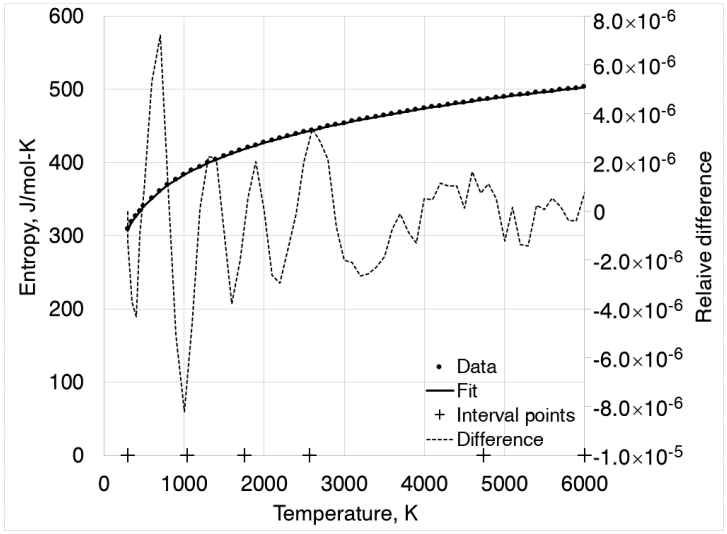


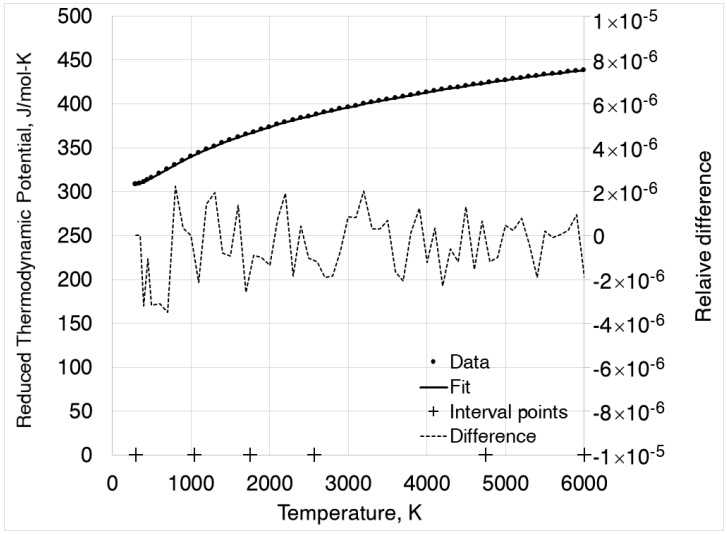
Titanium Bromide ($TiBr_2$) – Br2Ti1(g)

MW = 207.69 g/mol – $\Delta_f H^o$(298.15 K) = -179.075 kJ/mol – S^o(298.15 K) = 308.685 J/mol-K – p^o = 0.1 MPa



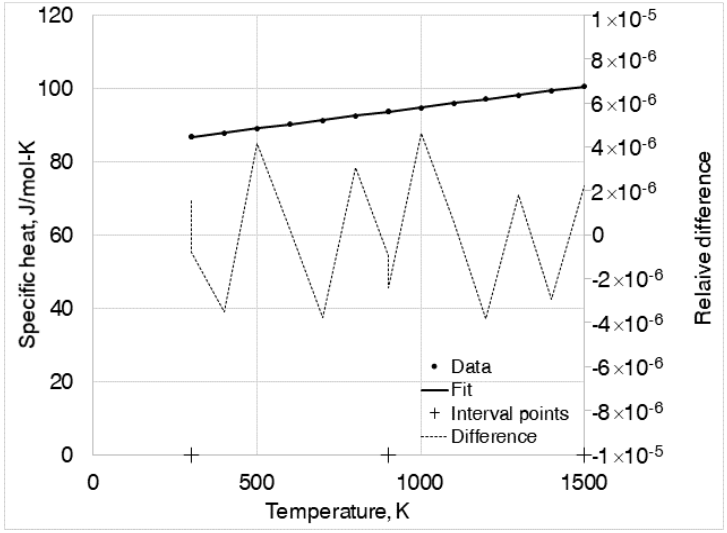


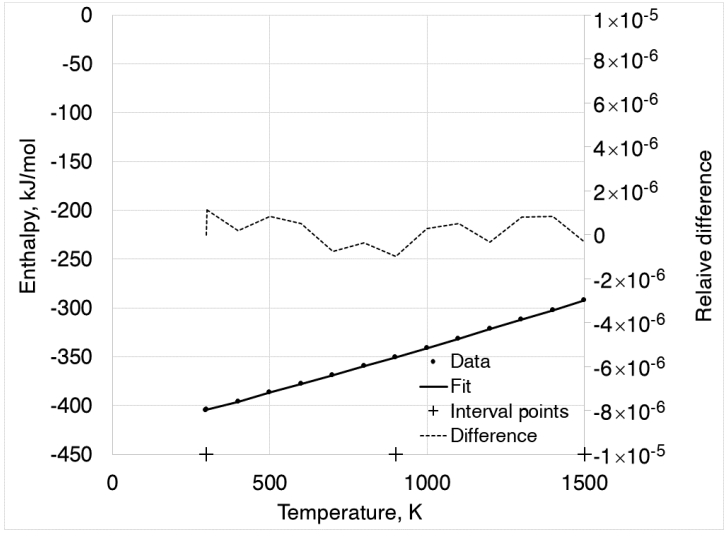


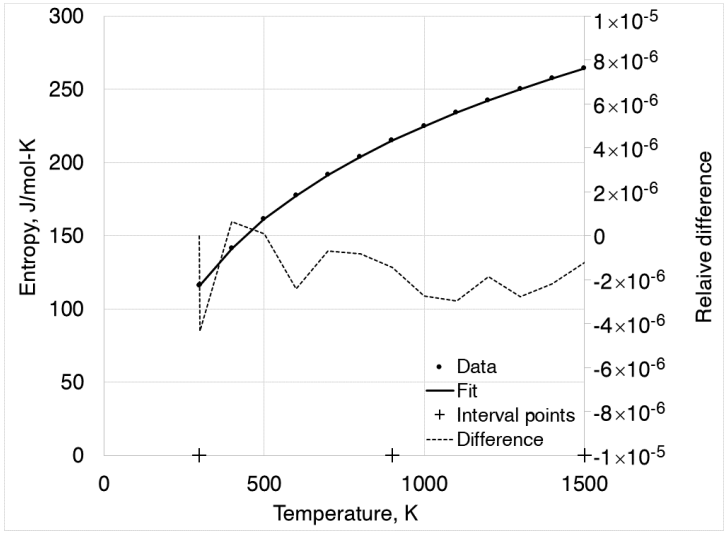


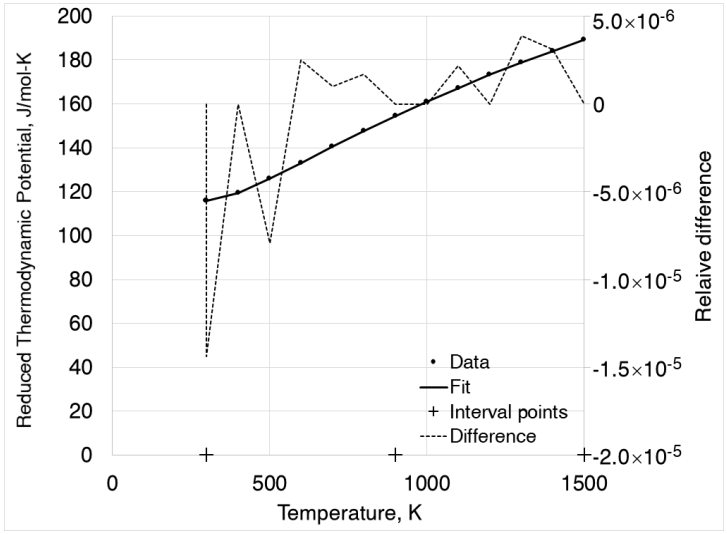
Zirconium Bromide ($ZrBr_2$) – Br2Zr1(cr)

MW = 251.03 g/mol – $\Delta_f H^o$(298.15 K) = -404.593 kJ/mol – S^o(298.15 K) = 115.897 J/mol-K – p^o = 0.1 MPa



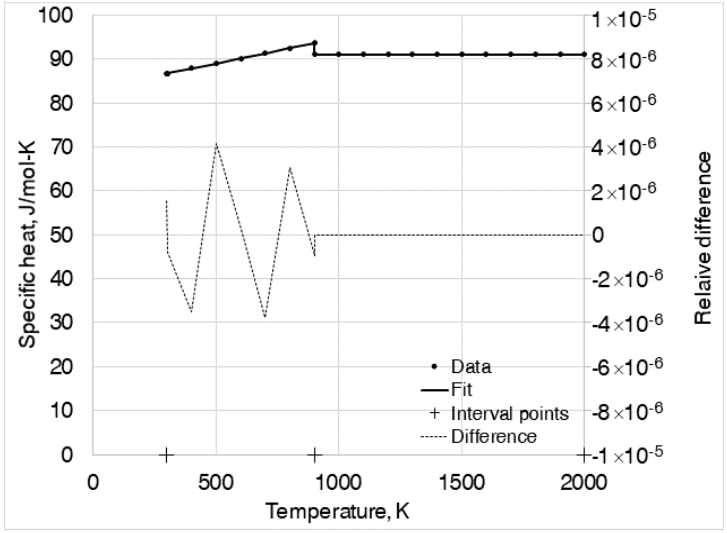


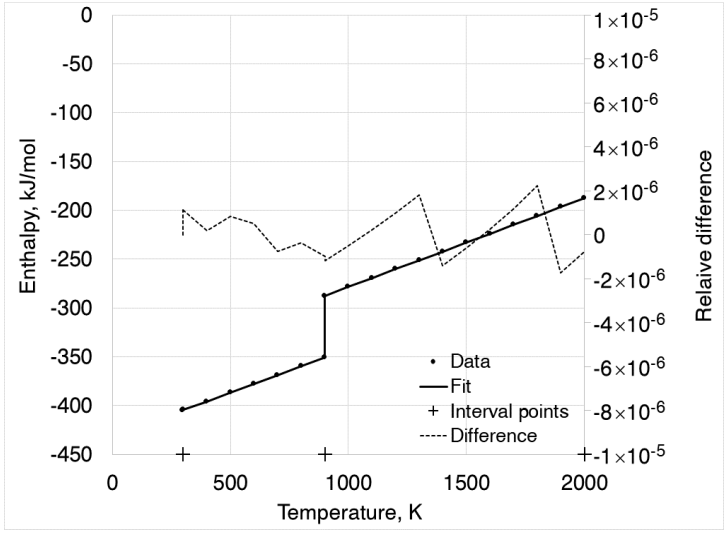


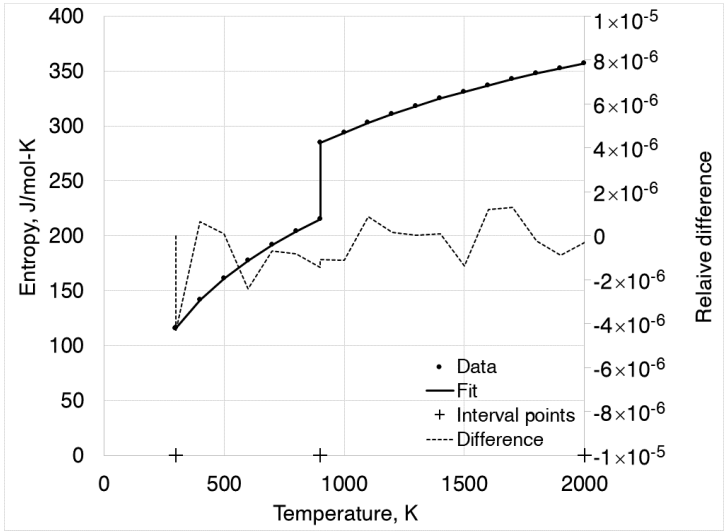


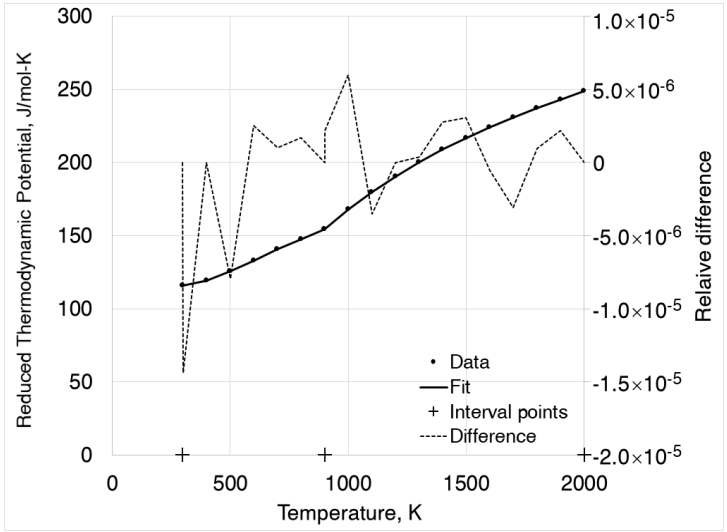
Zirconium Bromide ($ZrBr_2$) – Br2Zr1(cr,l)

MW = 251.03 g/mol – $\Delta_f H^o$(298.15 K) = -404.593 kJ/mol – S^o(298.15 K) = 115.897 J/mol-K – p^o = 0.1 MPa



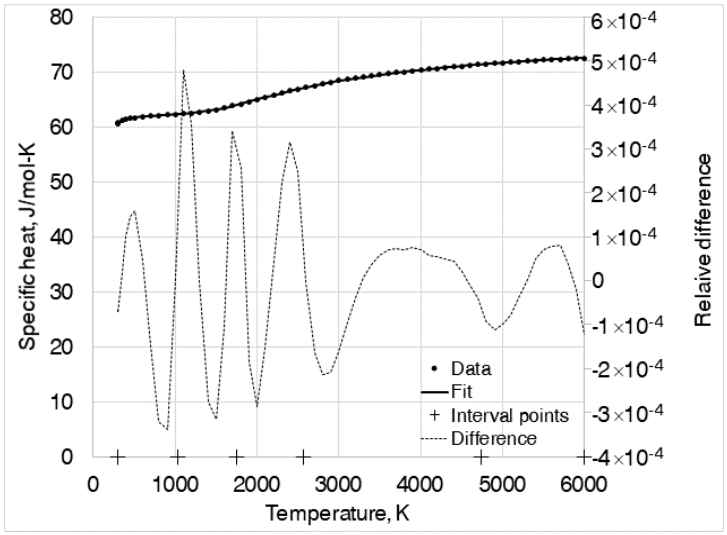


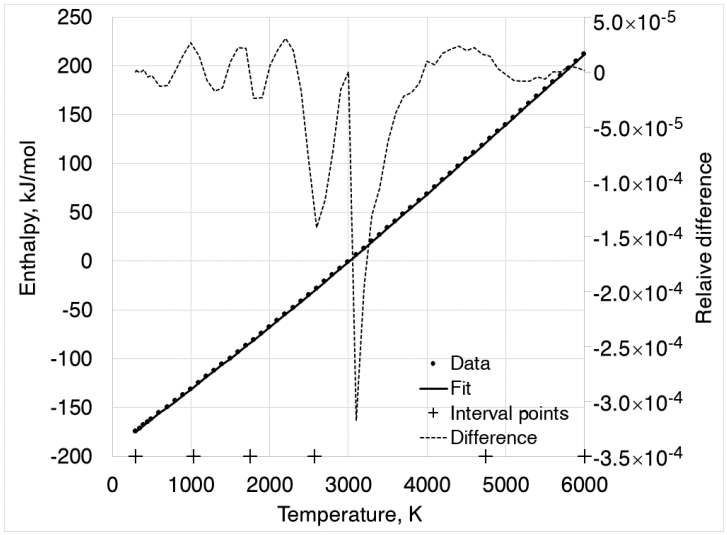


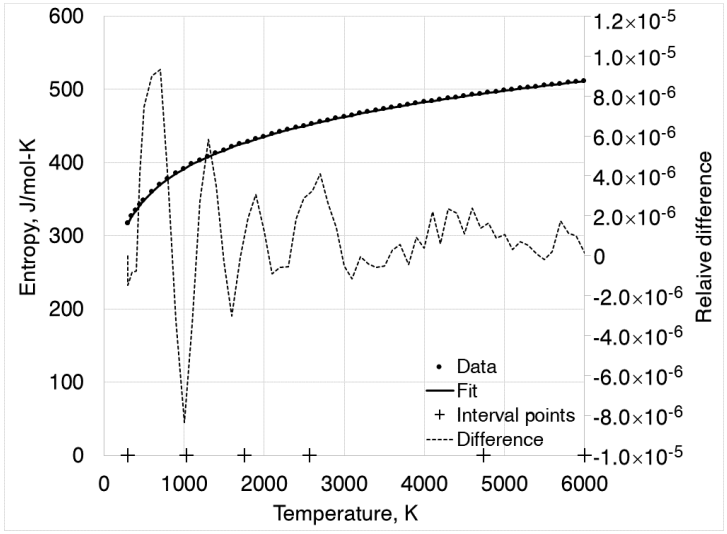


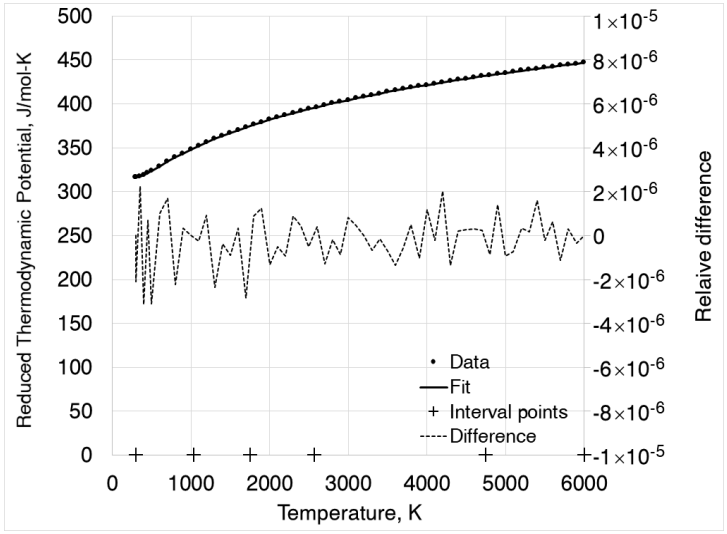
Zirconium Bromide ($ZrBr_2$) – Br2Zr1(g)

MW = 251.03 g/mol – $\Delta_f H^o$(298.15 K) = -174.473 kJ/mol – S^o(298.15 K) = 316.777 J/mol-K – p^o = 0.1 MPa



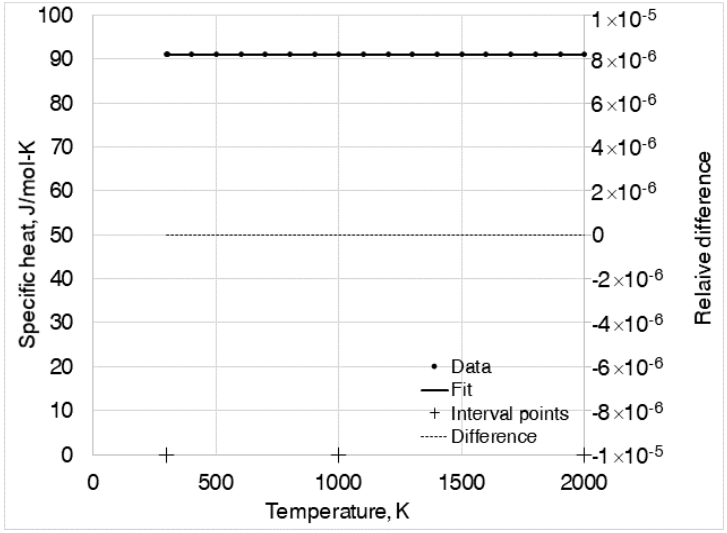


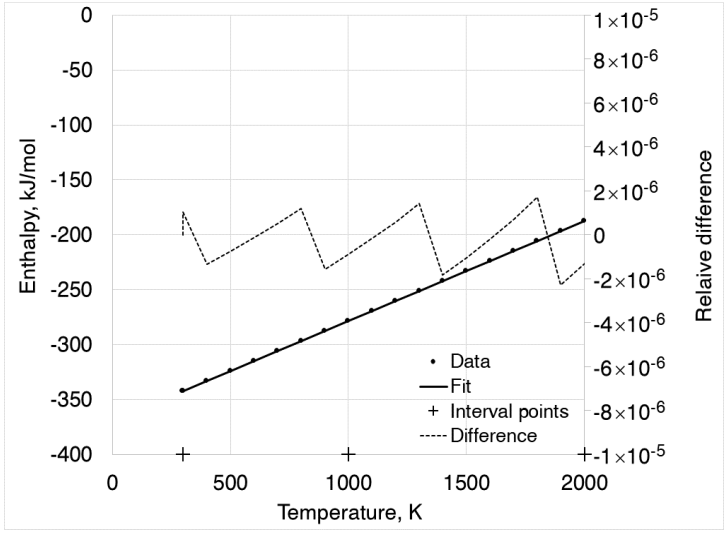


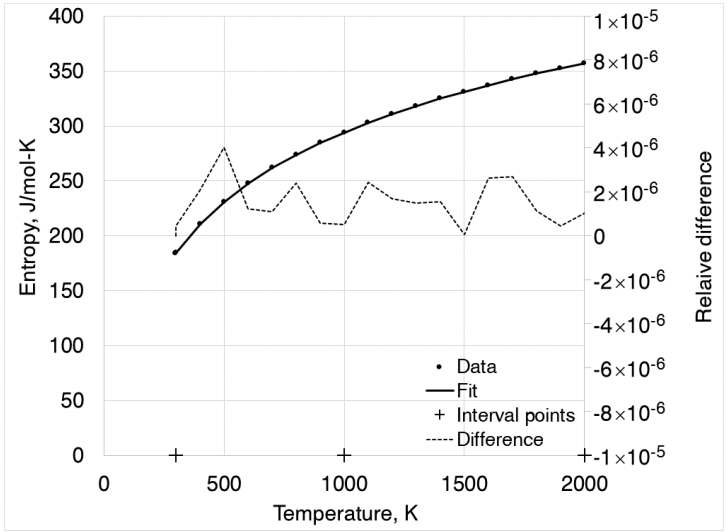


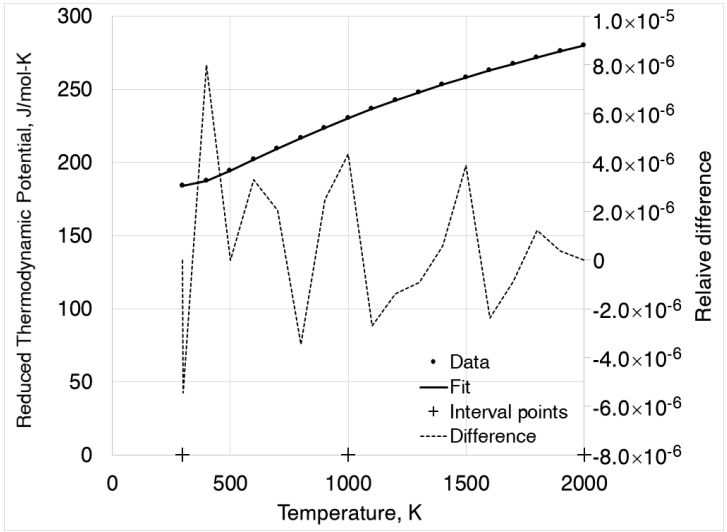
Zirconium Bromide ($ZrBr_2$) – Br2Zr1(l)

MW = 251.03 g/mol – $\Delta_f H^o$(298.15 K) = -342.333 kJ/mol – S^o(298.15 K) = 184.024 J/mol-K – p^o = 0.1 MPa



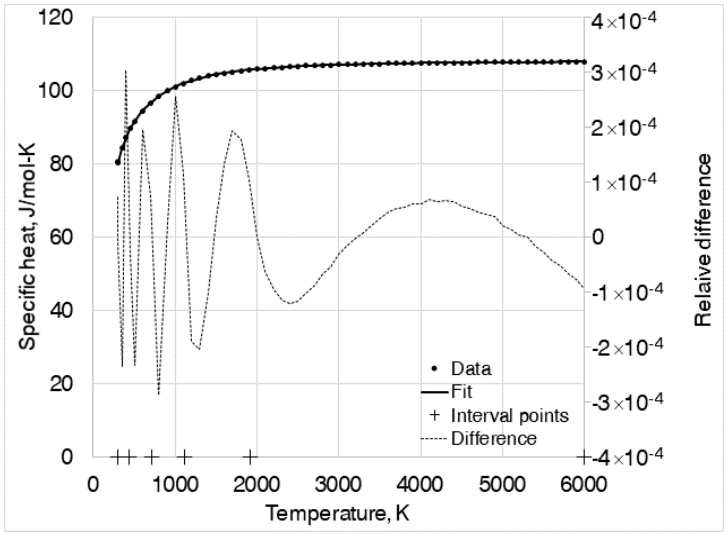


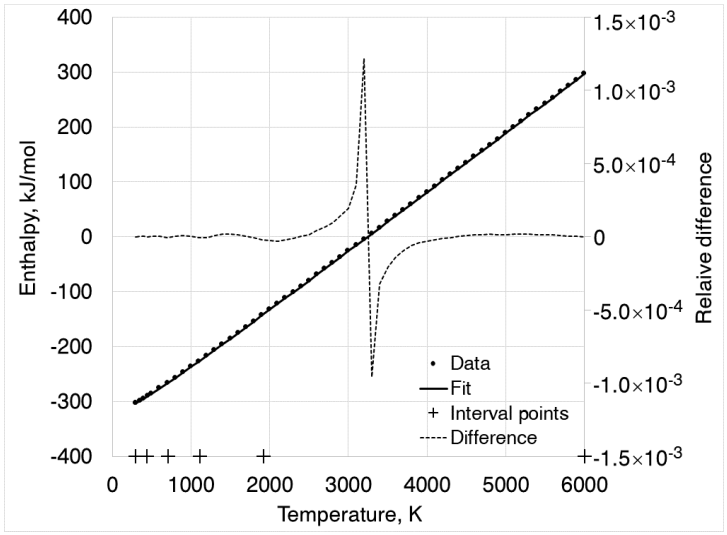


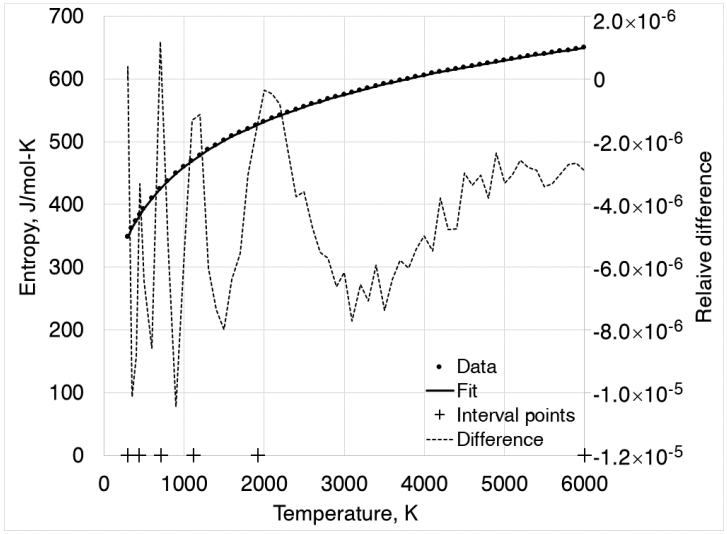


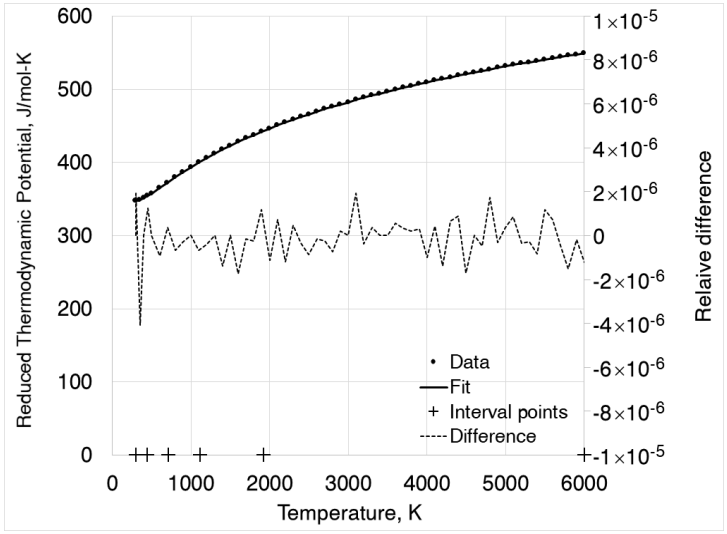
Tribromosilane ($SiHBr_3$) – Br3H1Si1(g)

MW = 268.805 g/mol – $\Delta_f H^o$(298.15 K) = -302.922 kJ/mol – S^o(298.15 K) = 348.049 J/mol-K – p^o = 0.1 MPa



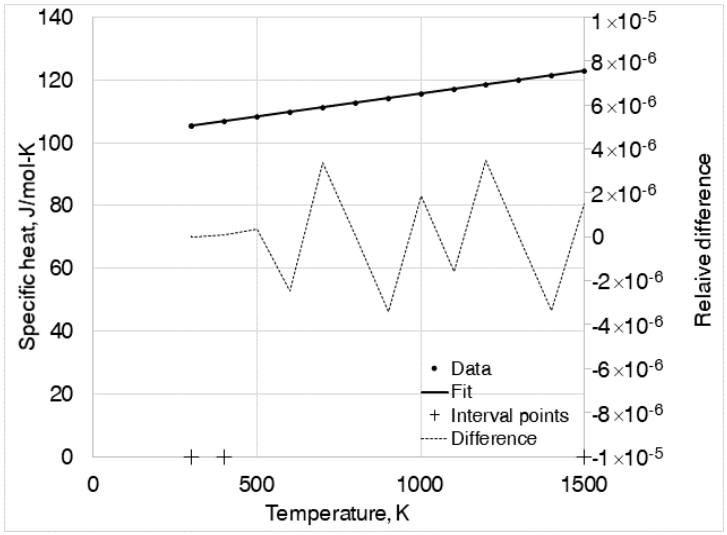


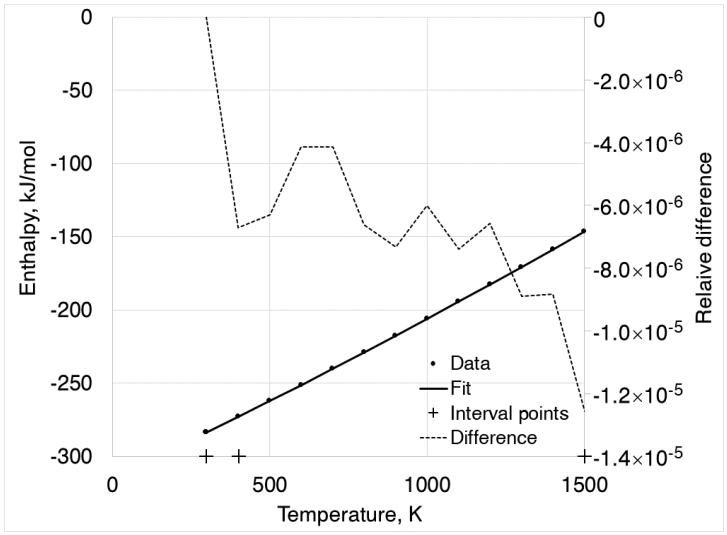


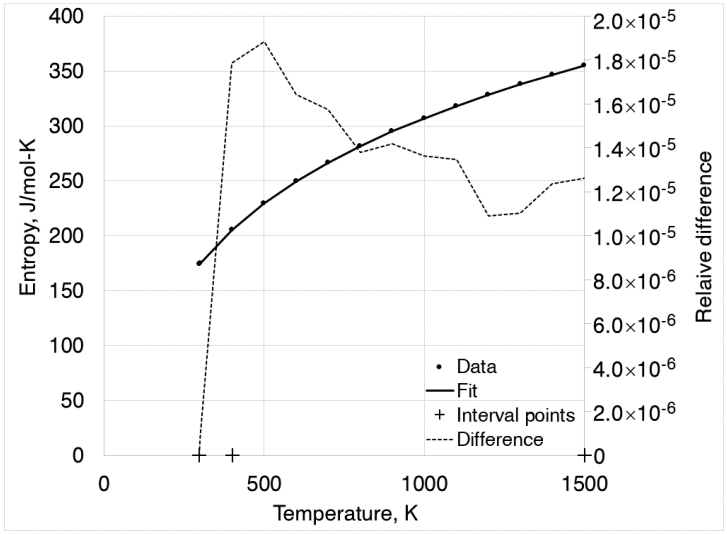


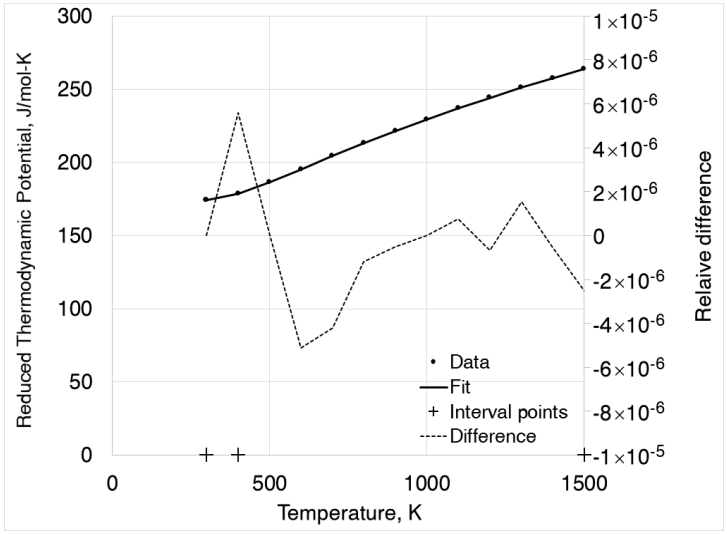
Molybdenum Bromide ($MoBr_3$) – Br3Mo1(cr)

MW = 335.65 g/mol – $\Delta_f H^o$(298.15 K) = -283.675 kJ/mol – S^o(298.15 K) = 174.473 J/mol-K – p^o = 0.1 MPa



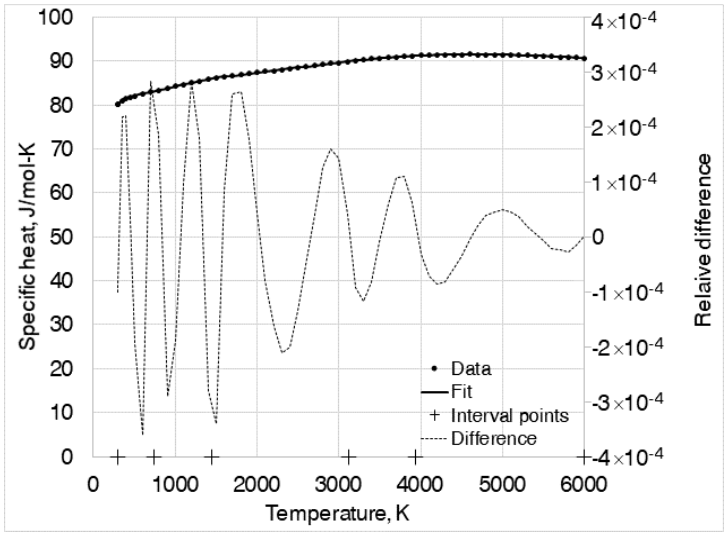


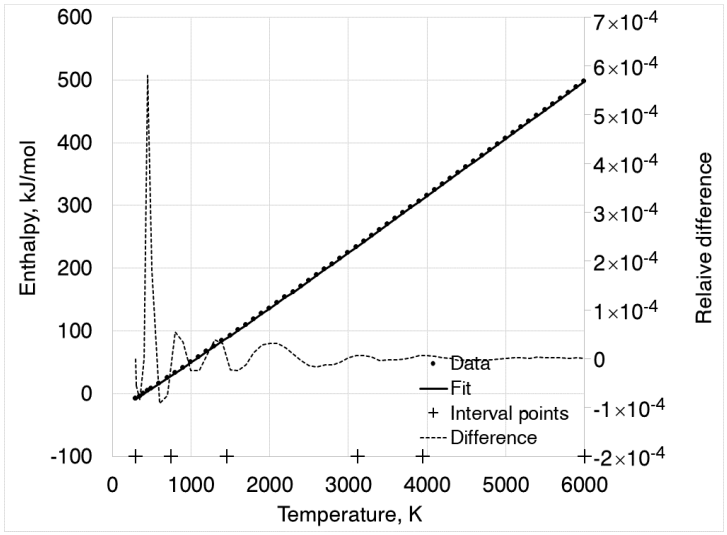


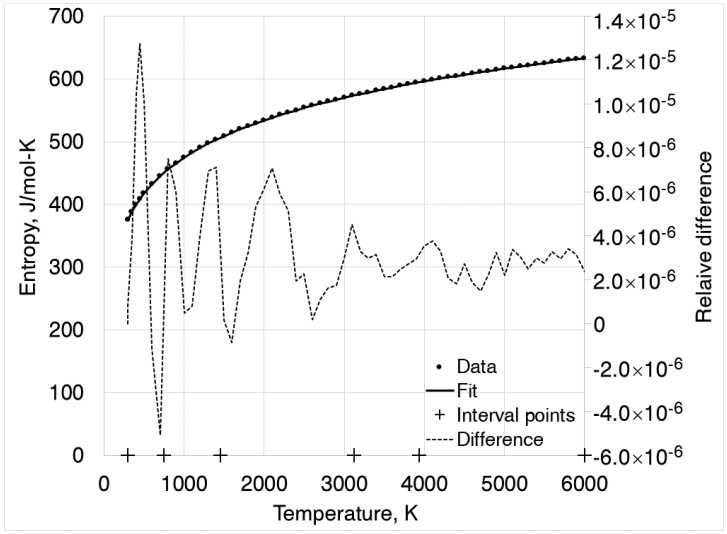


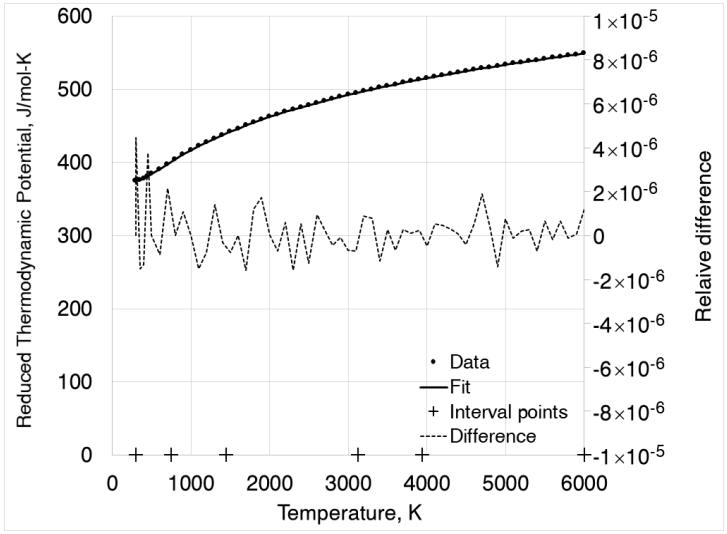
Molybdenum Bromide ($MoBr_3$) – Br3Mo1(g)

MW = 335.65 g/mol – $\Delta_f H^o$(298.15 K) = -8.368 kJ/mol – S^o(298.15 K) = 374.99 J/mol-K – p^o = 0.1 MPa



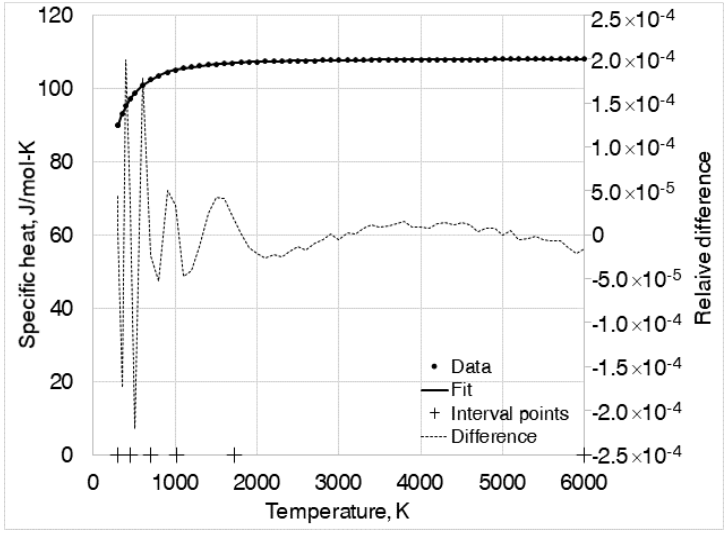


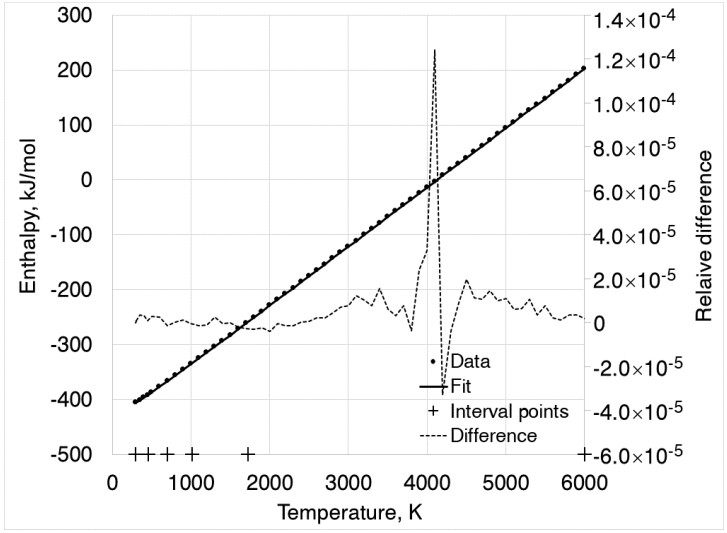


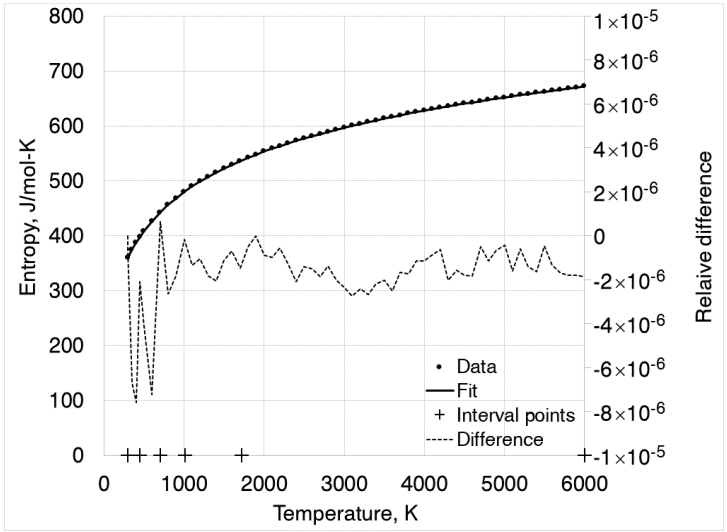


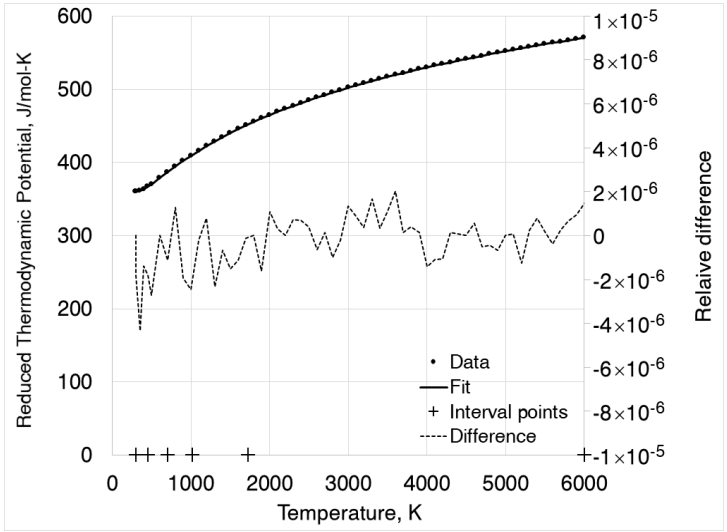
Phosphoryl Bromide ($POBr_3$) – Br3O1P1(g)

MW = 286.685 g/mol – $\Delta_f H^o$(298.15 K) = -406.022 kJ/mol – S^o(298.15 K) = 359.832 J/mol-K – p^o = 0.1 MPa



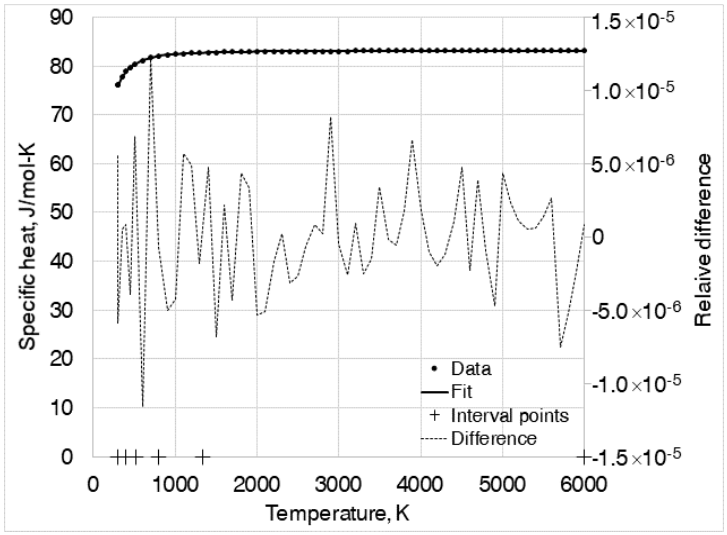


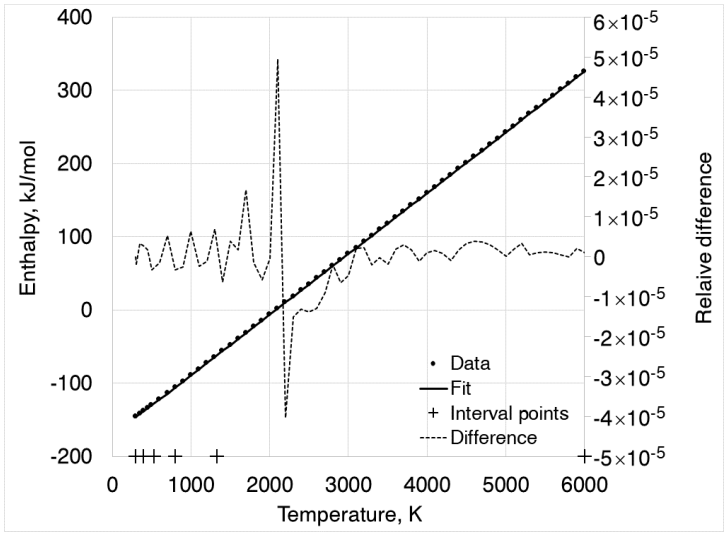


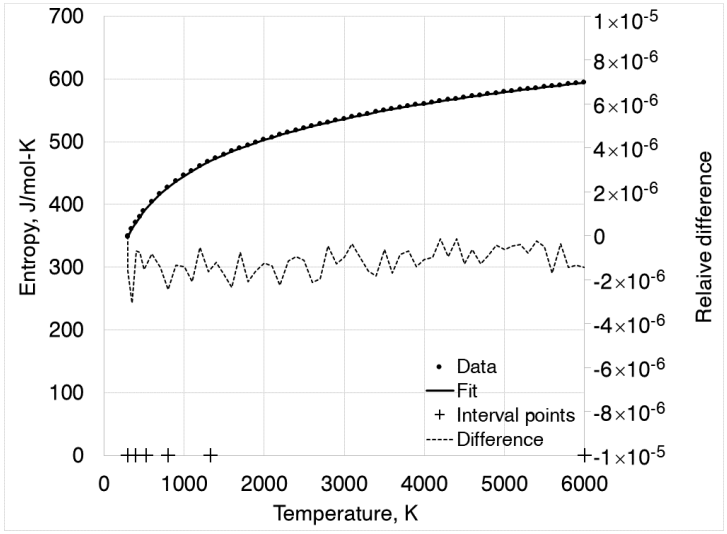


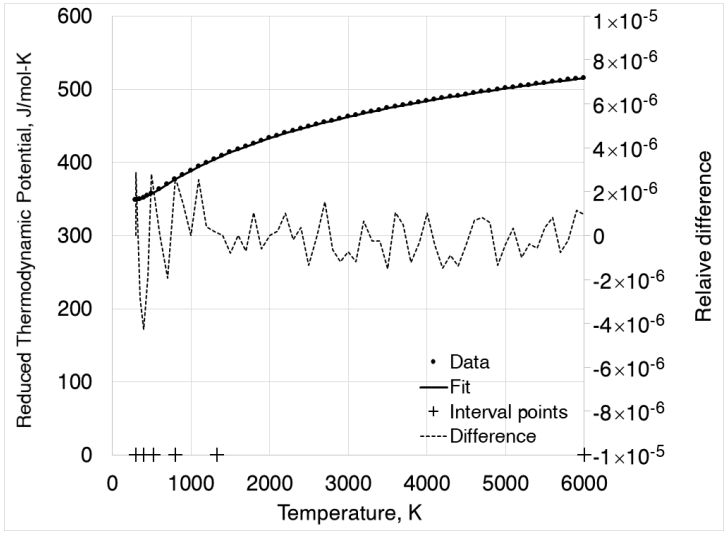
Phosphorus Bromide (PBr_3) – Br3P1(g)

MW = 270.686 g/mol – $\Delta_f H^o$(298.15 K) = -146.022 kJ/mol – S^o(298.15 K) = 348.234 J/mol-K – p^o = 0.1 MPa



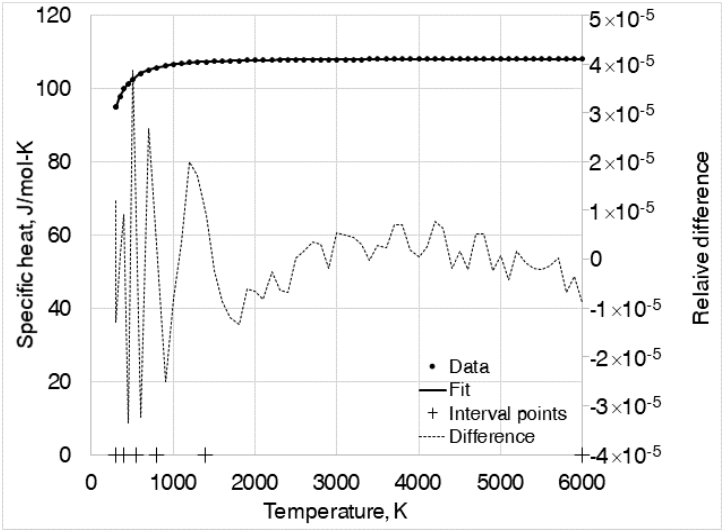


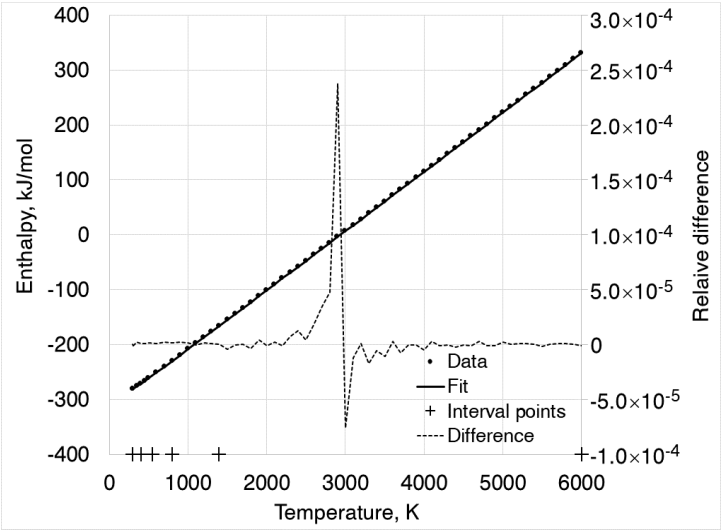


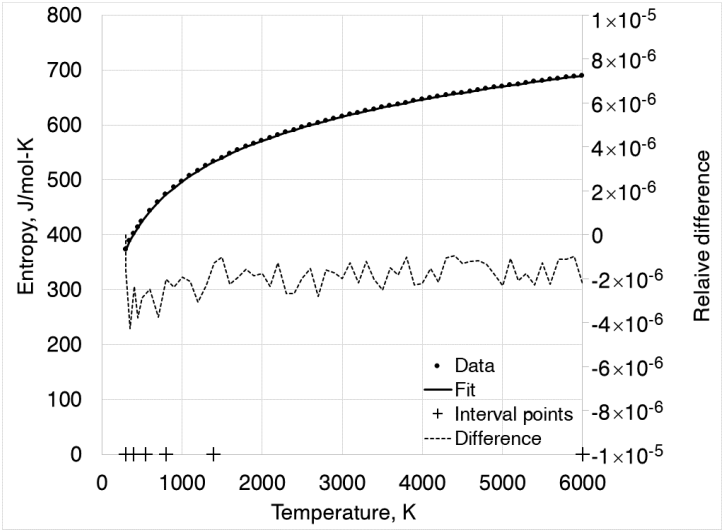


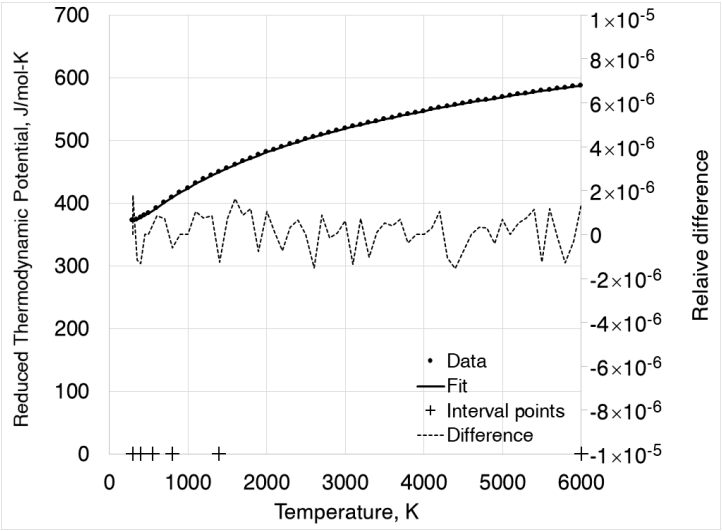
Thiophosphoryl Bromide ($PSBr_3$) – Br3P1S1(g)

MW = 302.75 g/mol – $\Delta_f H^o$(298.15 K) = -281.052 kJ/mol – S^o(298.15 K) = 372.828 J/mol-K – p^o = 0.1 MPa



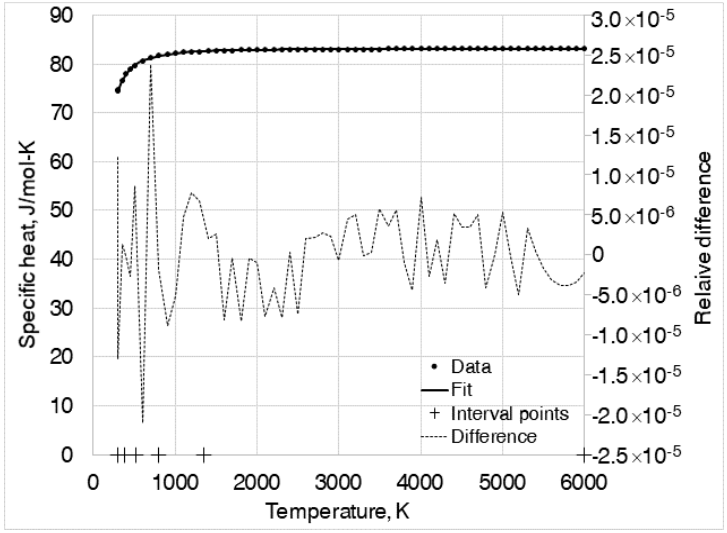


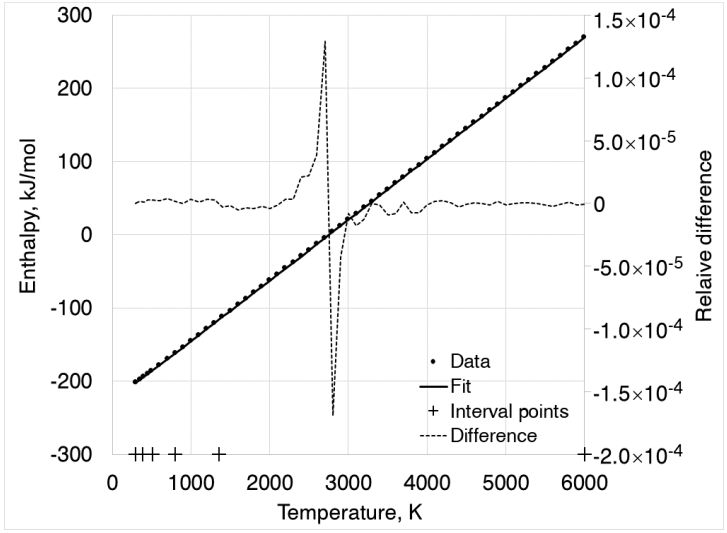


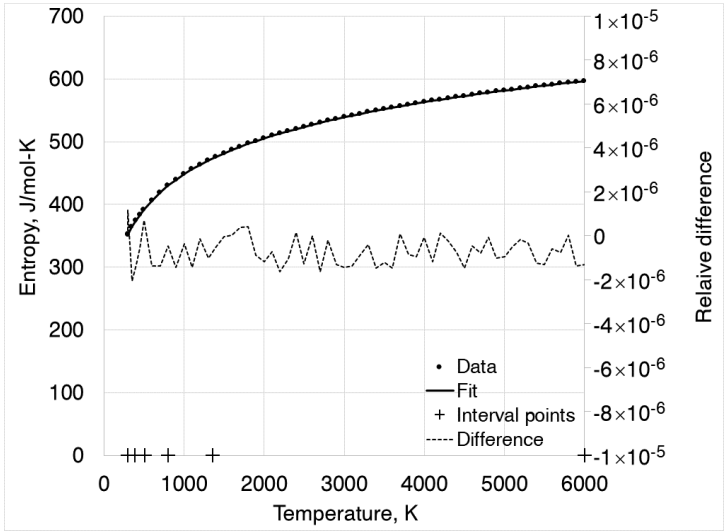


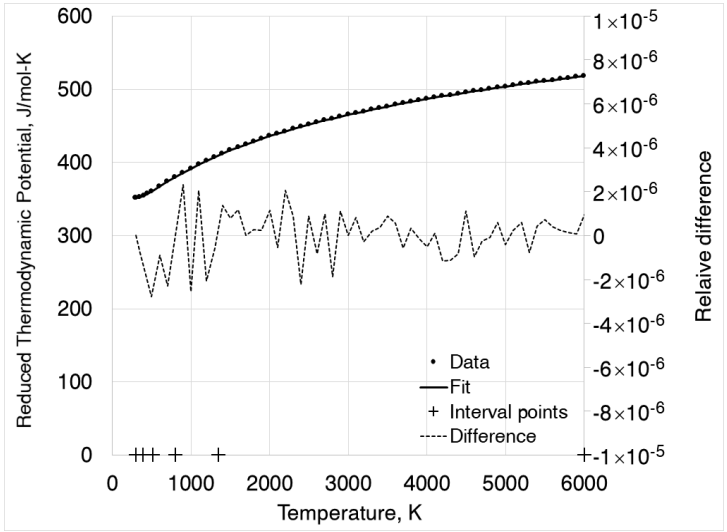
Tribromosilyl ($SiBr_3$) – Br3Si1(g)

MW = 267.798 g/mol – $\Delta_f H^o$(298.15 K) = -201.694 kJ/mol – S^o(298.15 K) = 351.774 J/mol-K – p^o = 0.1 MPa



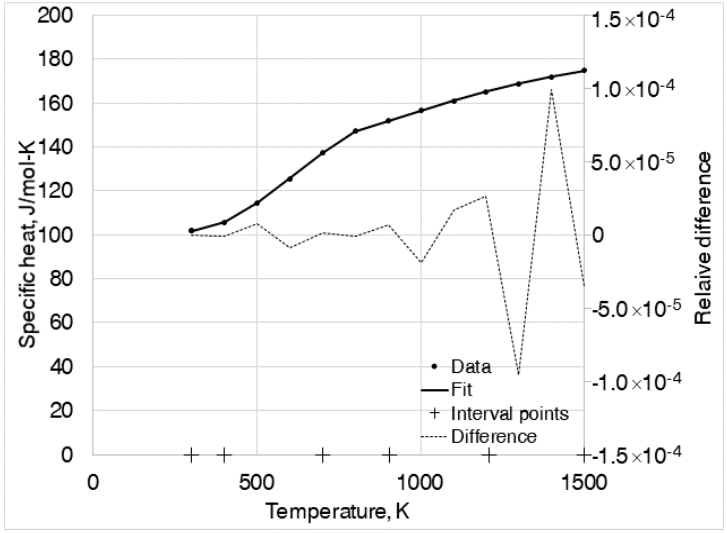


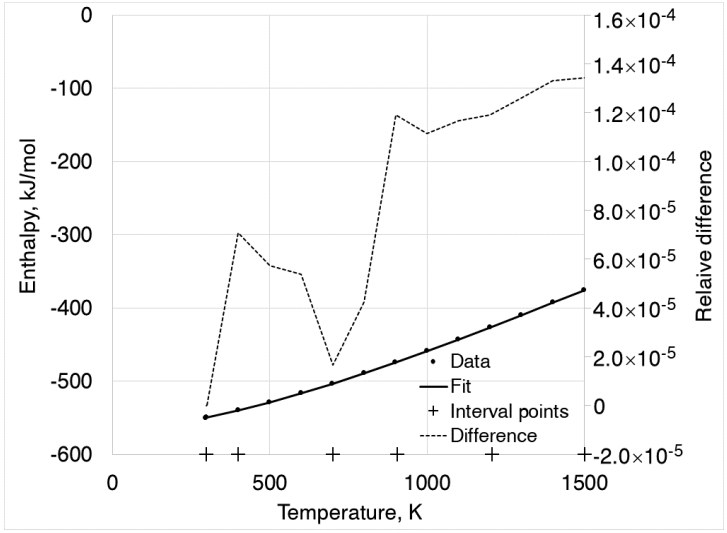


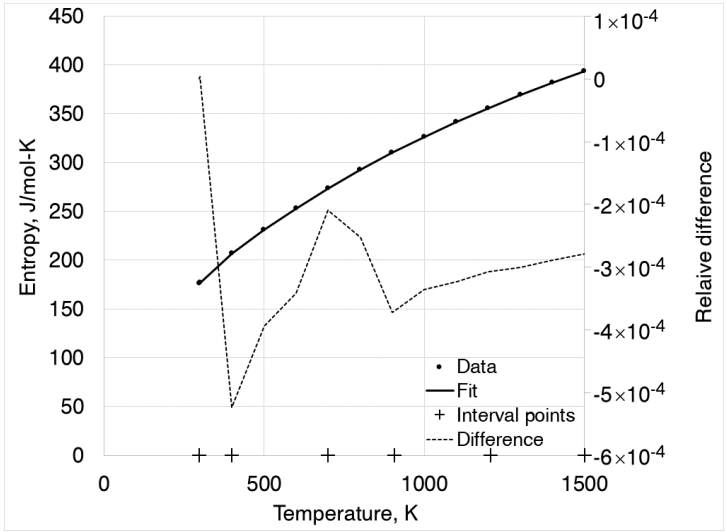


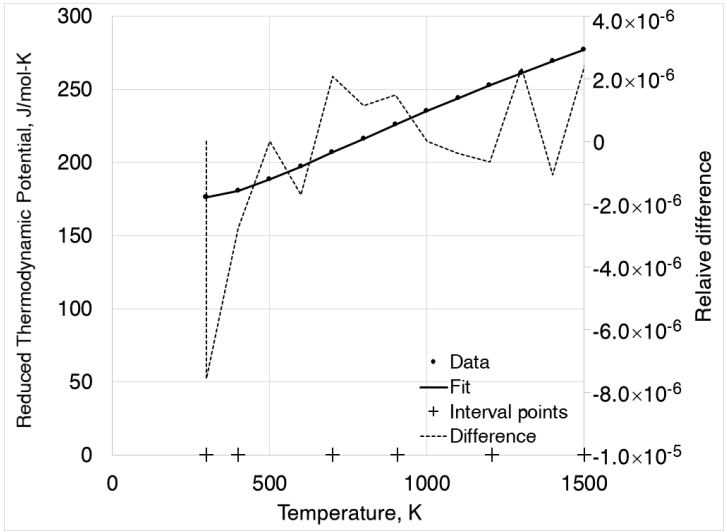
Titanium Bromide ($TiBr_3$) – Br3Ti1(cr)

MW = 287.59 g/mol – $\Delta_f H^o$(298.15 K) = -550.196 kJ/mol – S^o(298.15 K) = 176.443 J/mol-K – p^o = 0.1 MPa



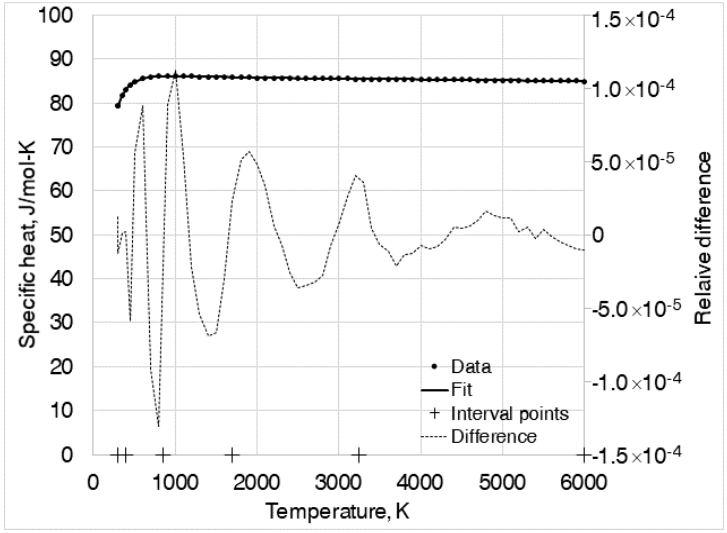


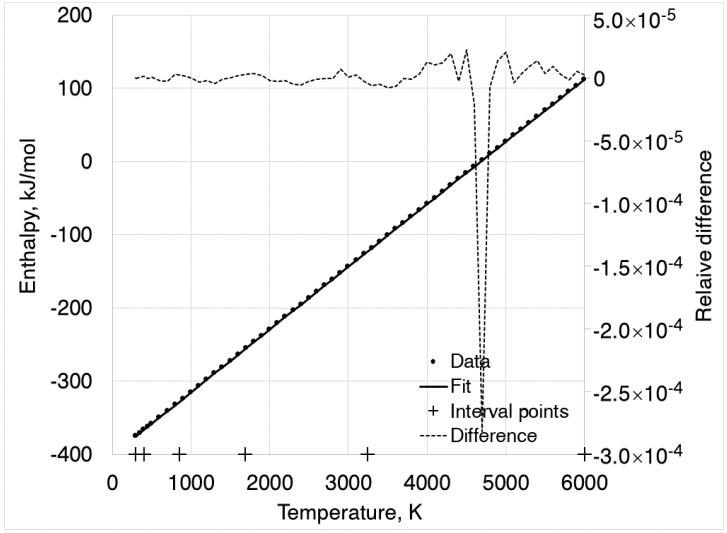


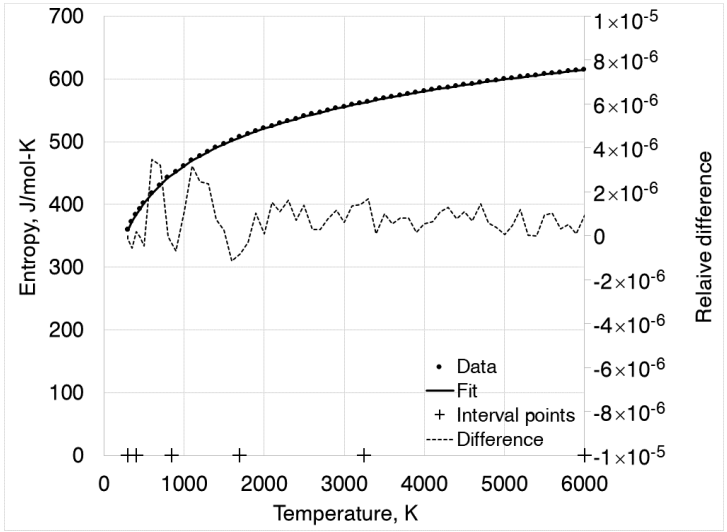


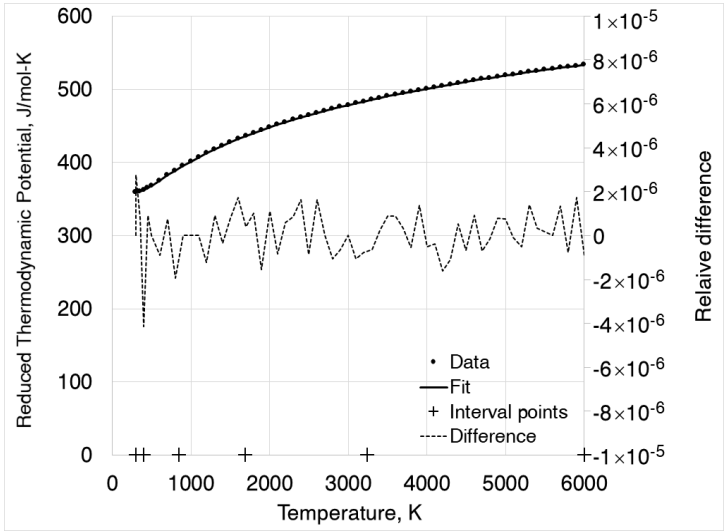
Titanium Bromide ($TiBr_3$) – Br3Ti1(g)

MW = 287.59 g/mol – $\Delta_f H^o$(298.15 K) = -374.886 kJ/mol – S^o(298.15 K) = 359.064 J/mol-K – p^o = 0.1 MPa



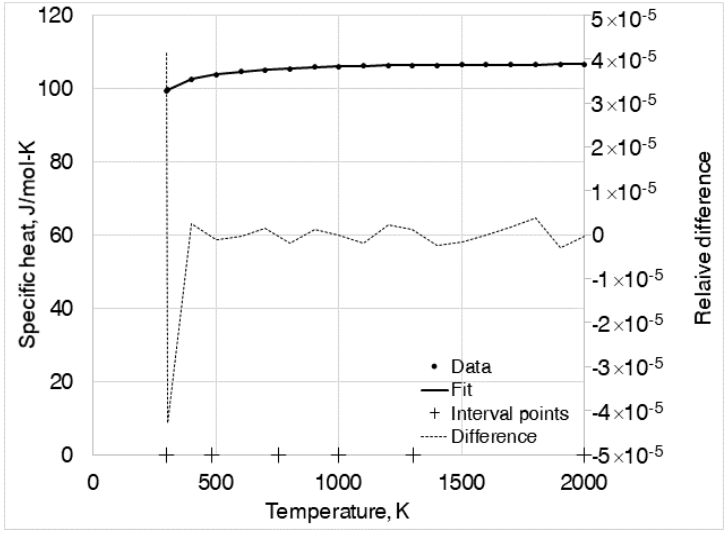


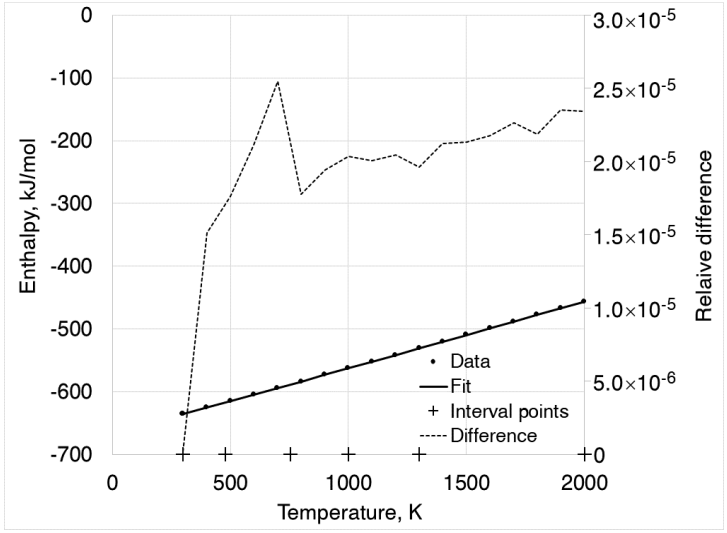


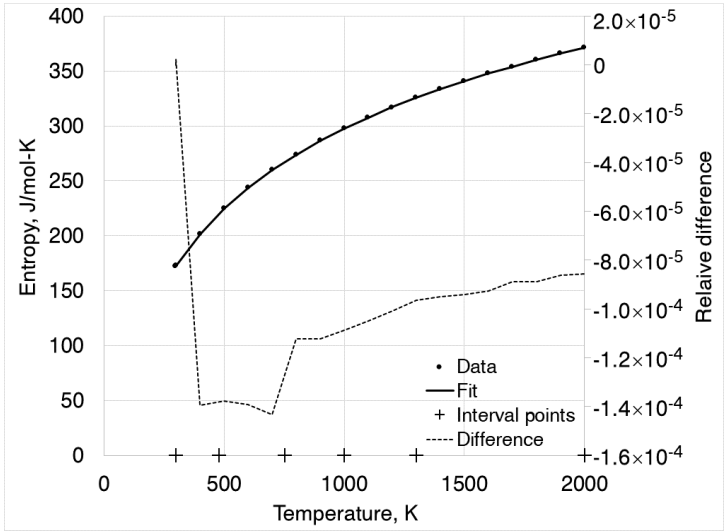


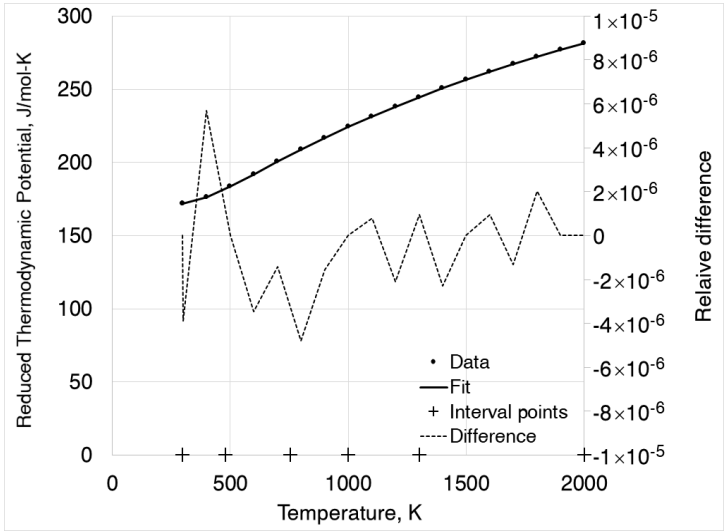
Zirconium Bromide ($ZrBr_3$) – Br3Zr1(cr)

MW = 330.93 g/mol – $\Delta_f H^o$(298.15 K) = -635.968 kJ/mol – S^o(298.15 K) = 172.054 J/mol-K – p^o = 0.1 MPa



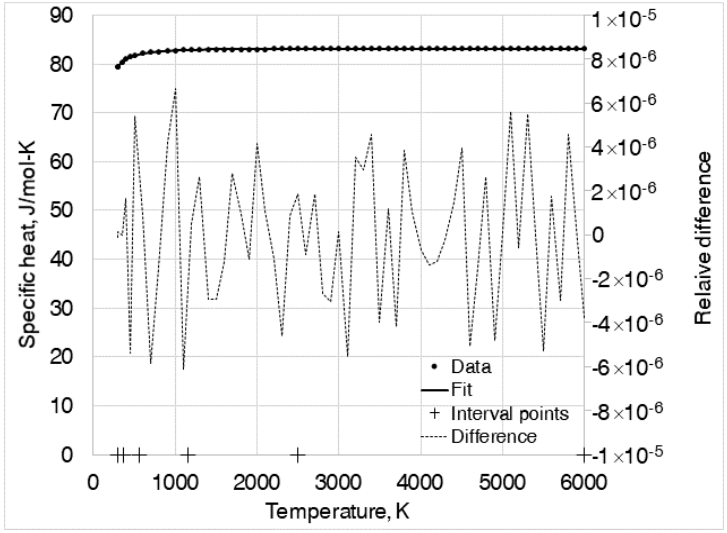


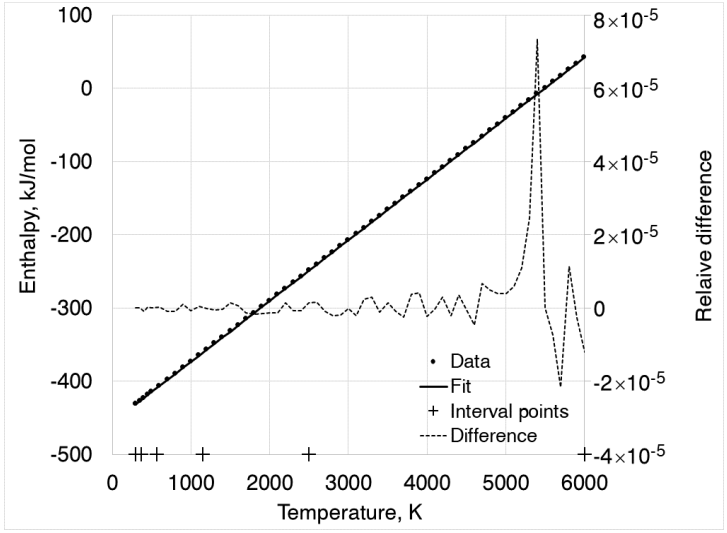


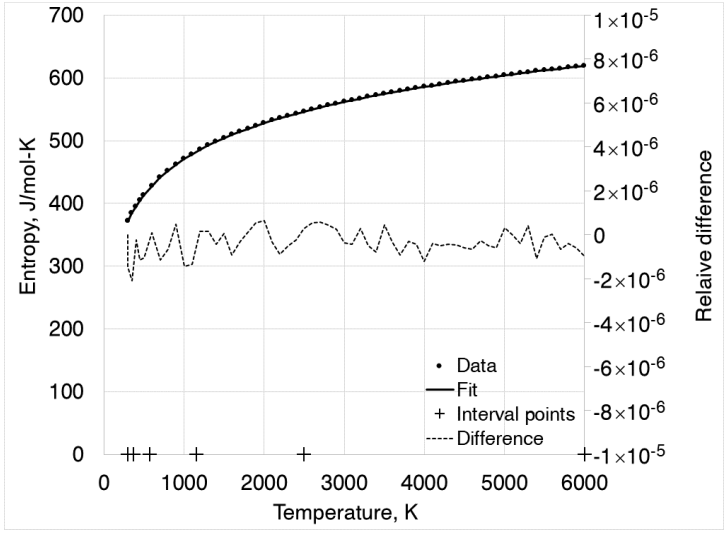


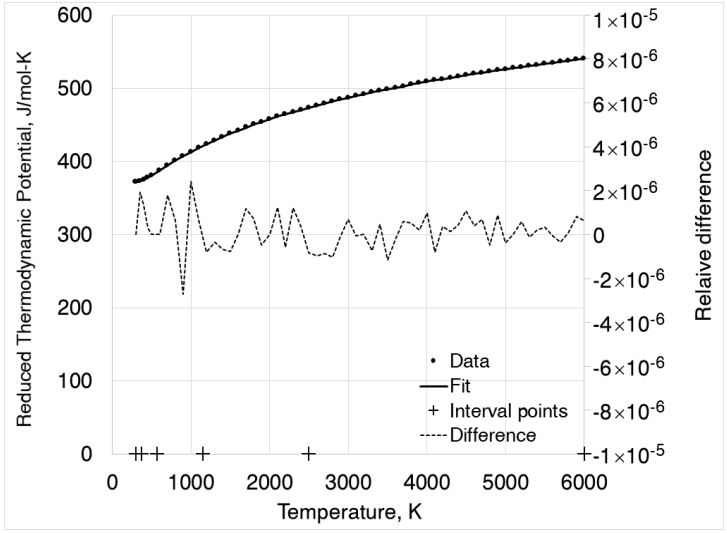
Zirconium Bromide ($ZrBr_3$) – Br3Zr1(g)

MW = 330.93 g/mol – $\Delta_f H^o$(298.15 K) = -430.952 kJ/mol – S^o(298.15 K) = 371.585 J/mol-K – p^o = 0.1 MPa



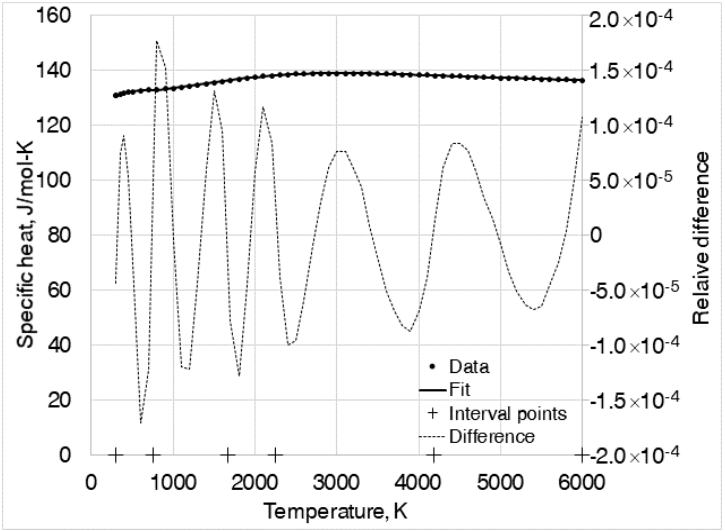


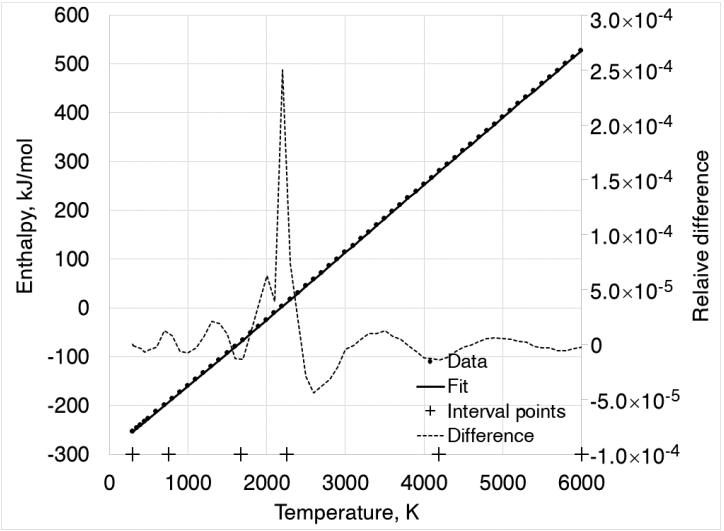


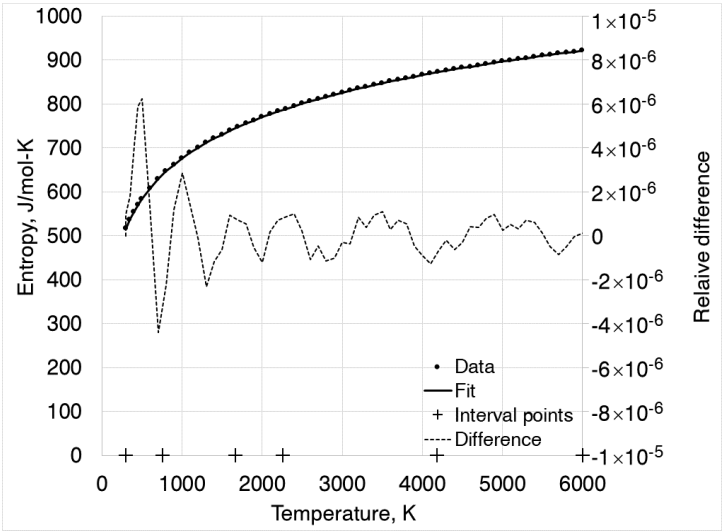


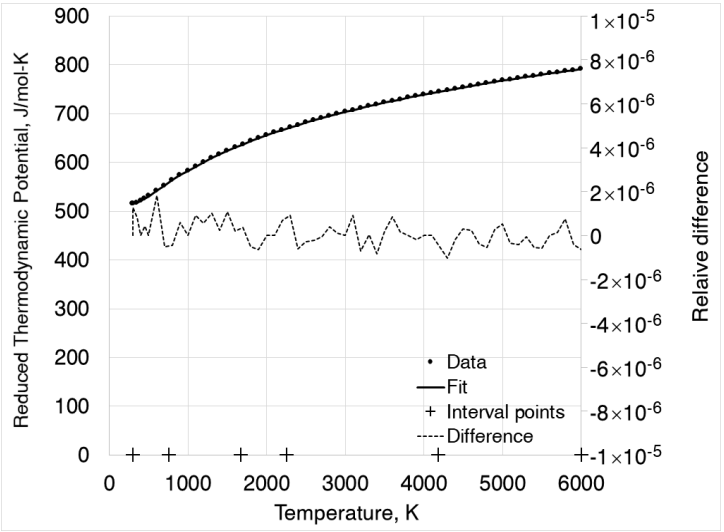
Iron Bromide ($(FeBr_2)_2$) – Br4Fe2(g)

MW = 431.31 g/mol – $\Delta_f H^o$(298.15 K) = -253.132 kJ/mol – S^o(298.15 K) = 516.018 J/mol-K – p^o = 0.1 MPa



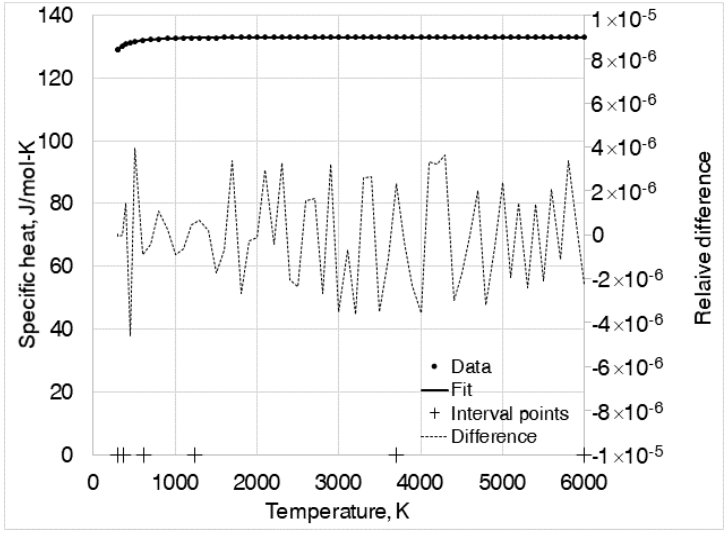


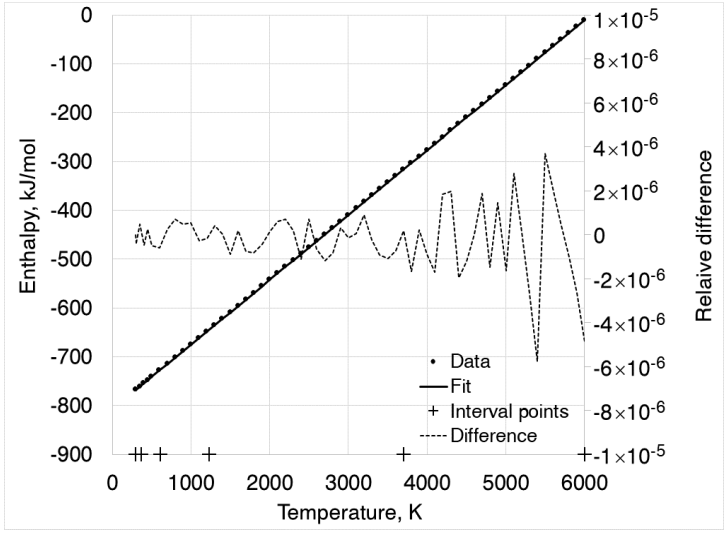


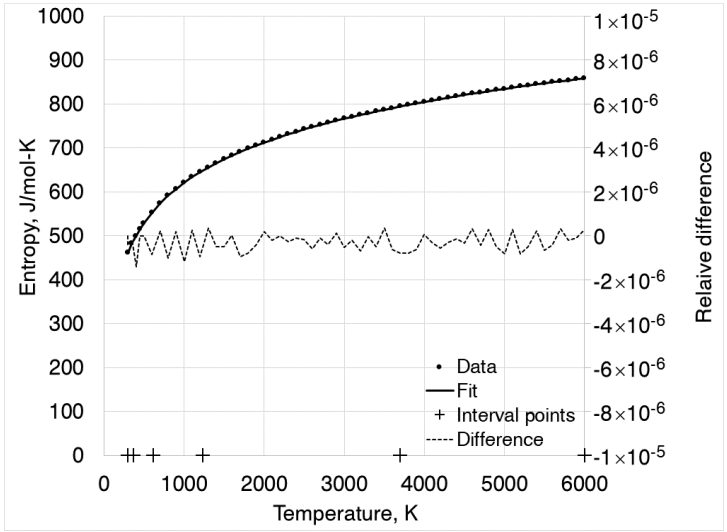


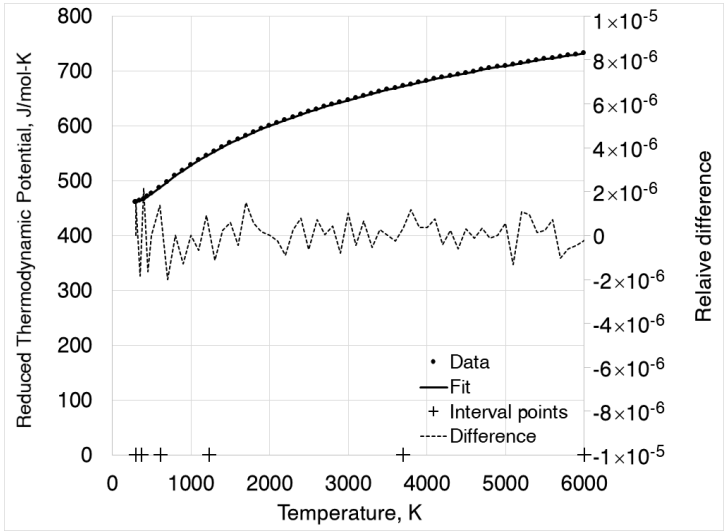
Magnesium Bromide ($(MgBr_2)_2$) – Br4Mg2(g)

MW = 368.226 g/mol – $\Delta_f H^o$(298.15 K) = -767.764 kJ/mol – S^o(298.15 K) = 461.347 J/mol-K – p^o = 0.1 MPa



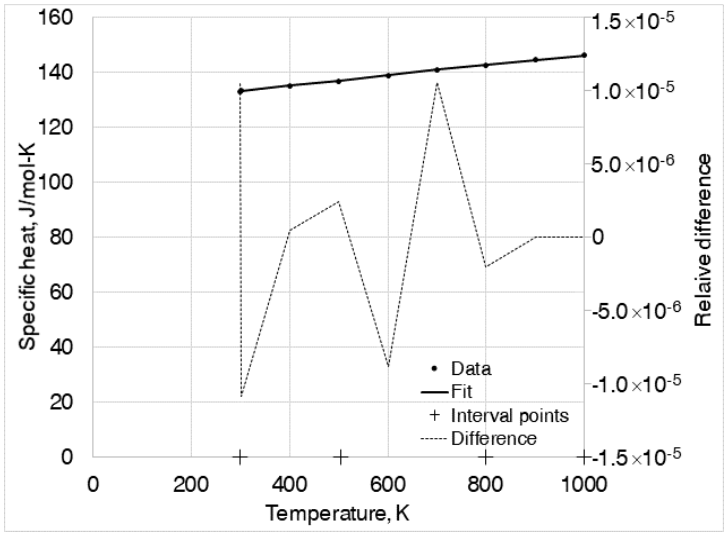


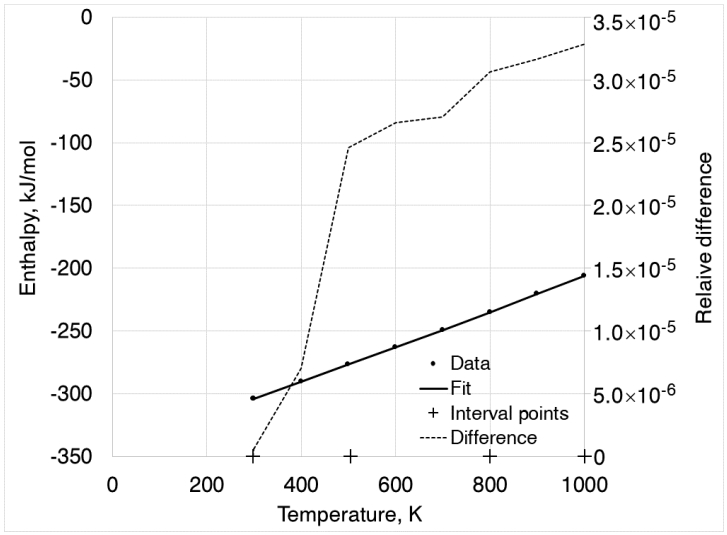


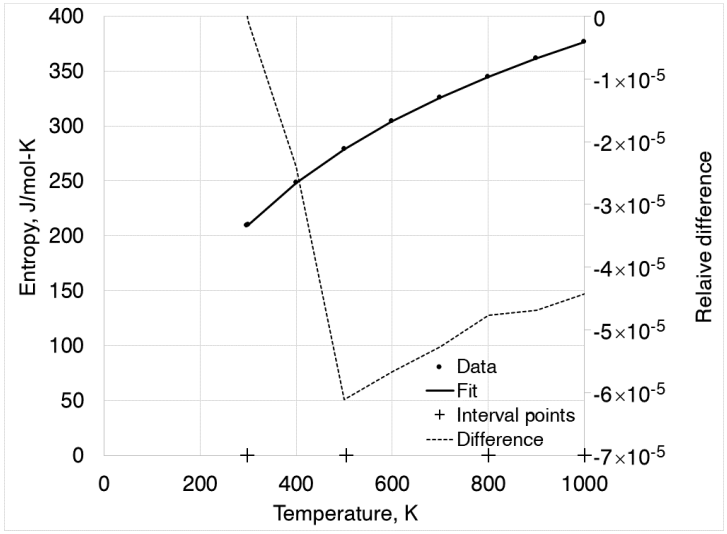


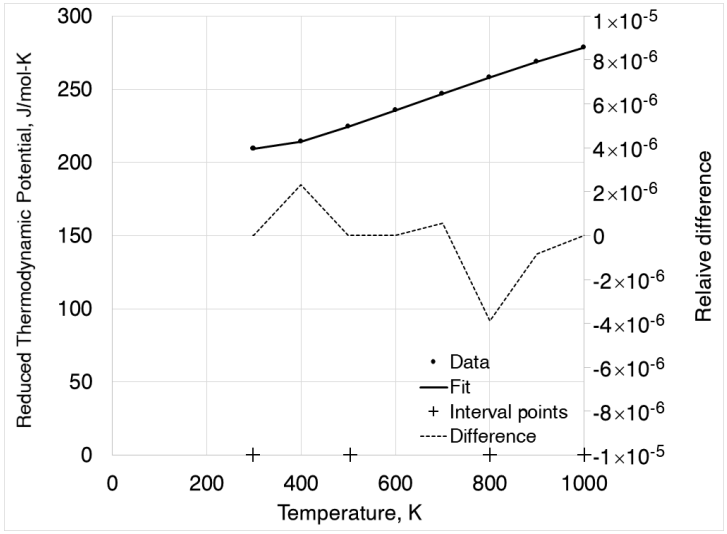
Molybdenum Bromide ($MoBr_4$) – Br4Mo1(cr)

MW = 415.56 g/mol – $\Delta_f H^o$(298.15 K) = -304.177 kJ/mol – S^o(298.15 K) = 209.2 J/mol-K – p^o = 0.1 MPa



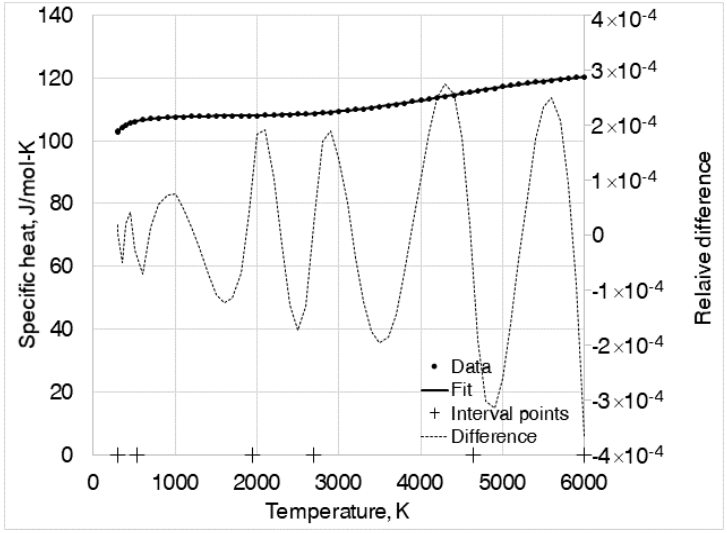


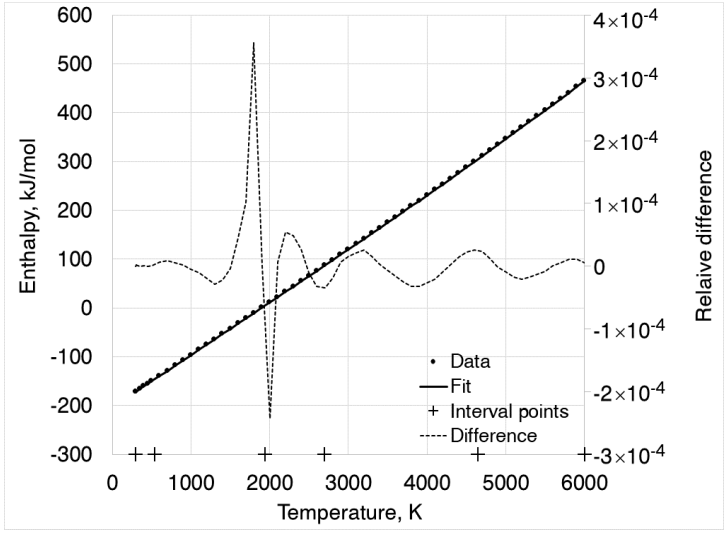


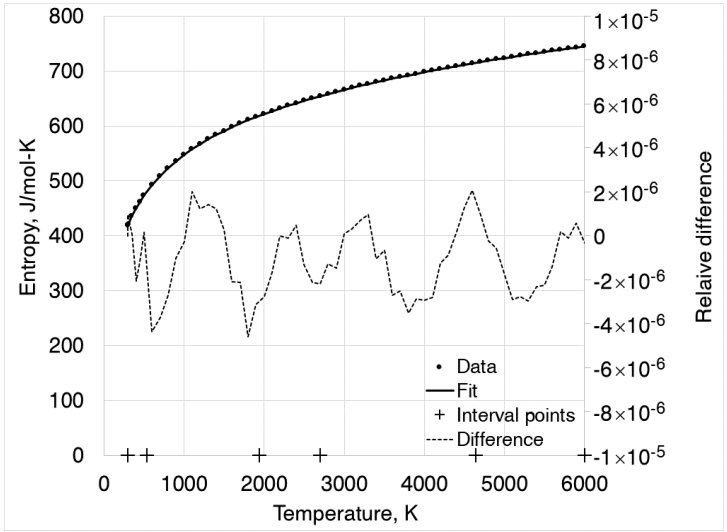


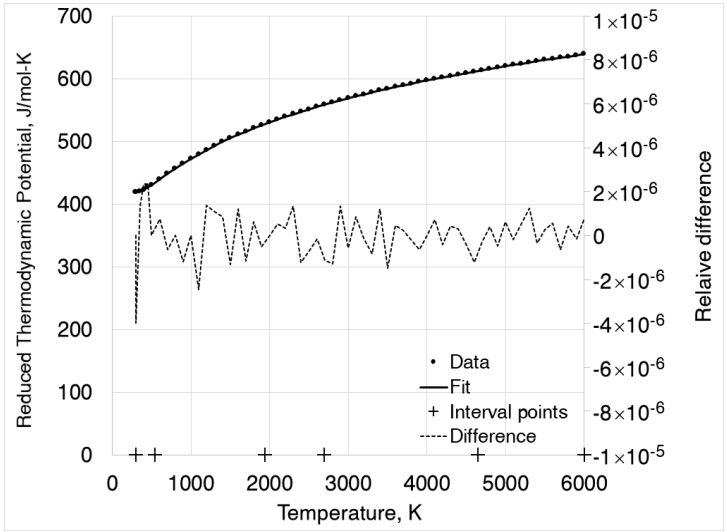
Molybdenum Bromide ($MoBr_4$) – Br4Mo1(g)

MW = 415.56 g/mol – $\Delta_f H^o$(298.15 K) = -171.126 kJ/mol – S^o(298.15 K) = 418.889 J/mol-K – p^o = 0.1 MPa



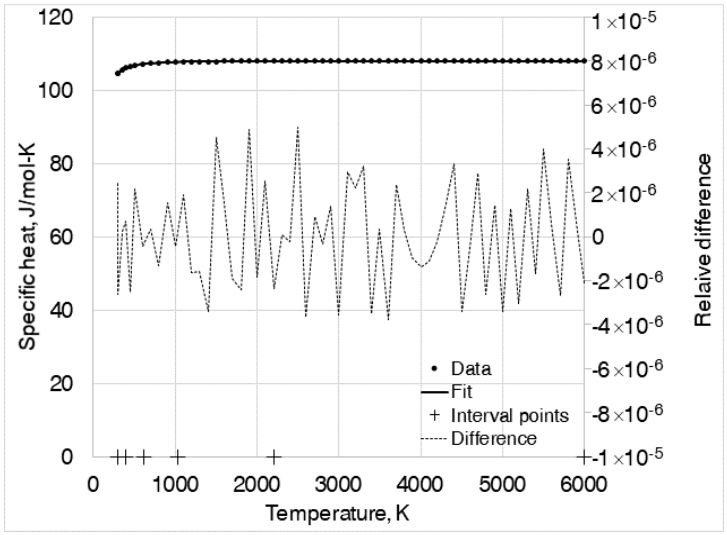


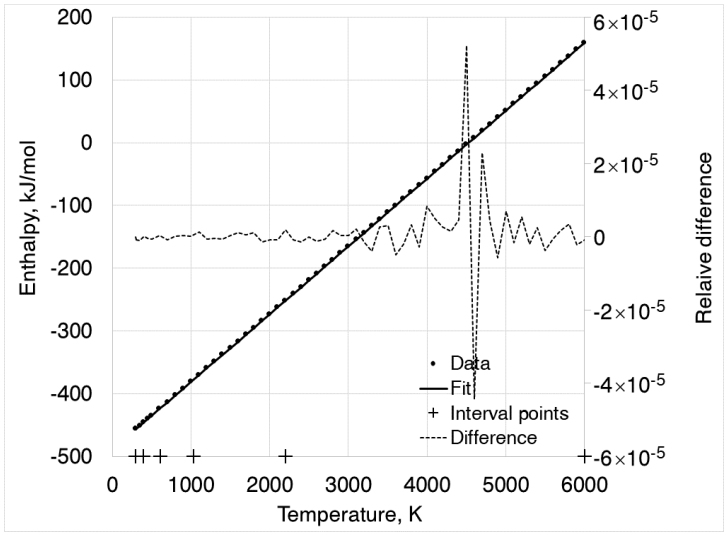


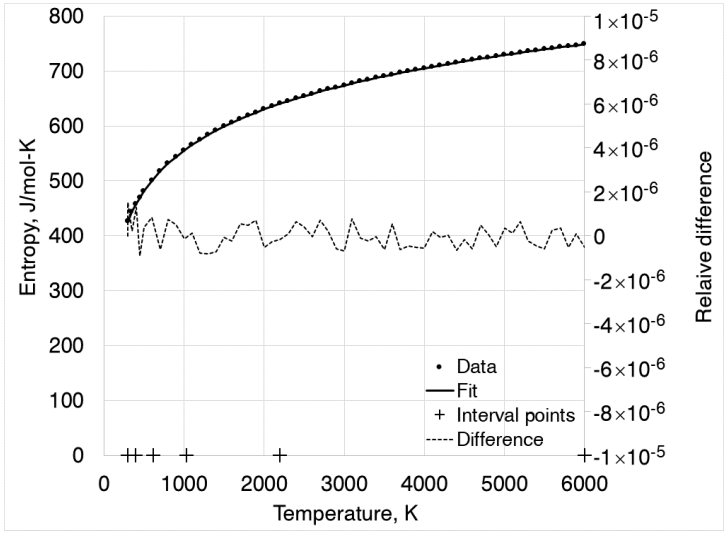


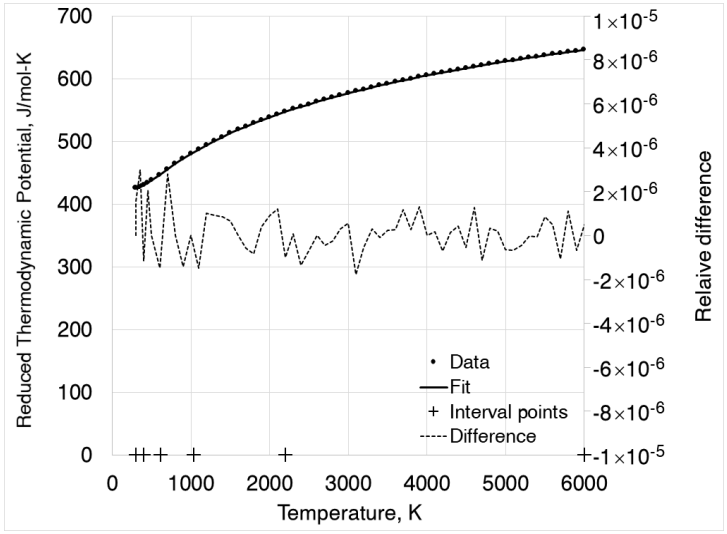
Lead Bromide ($PbBr_4$) – Br4Pb1(g)

MW = 526.8 g/mol – $\Delta_f H^o$(298.15 K) = -456.361 kJ/mol – S^o(298.15 K) = 426.201 J/mol-K – p^o = 0.1 MPa



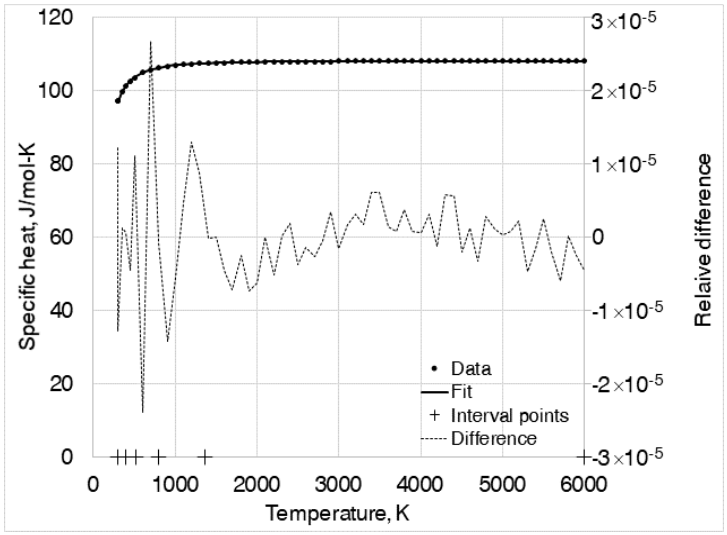


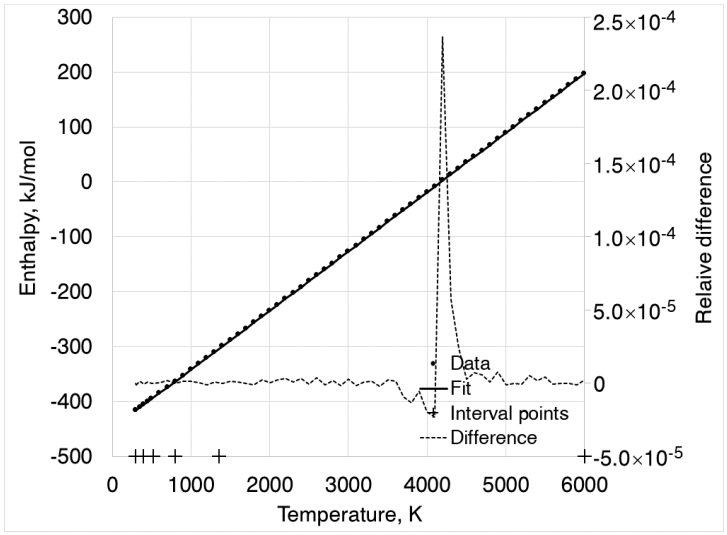


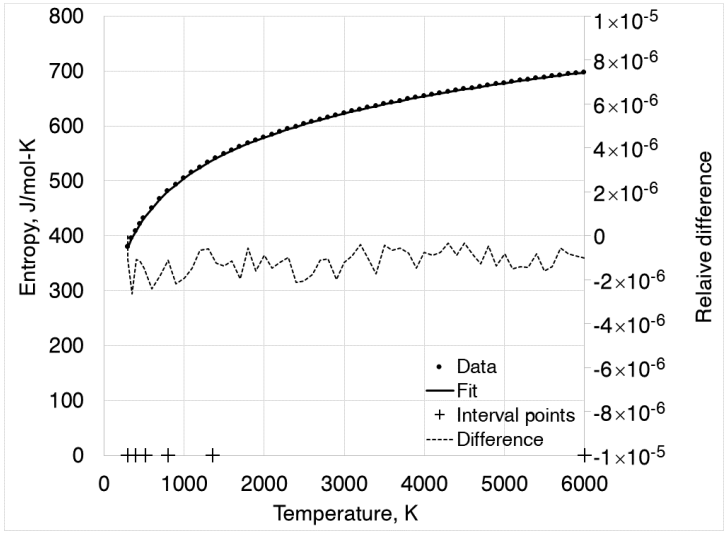


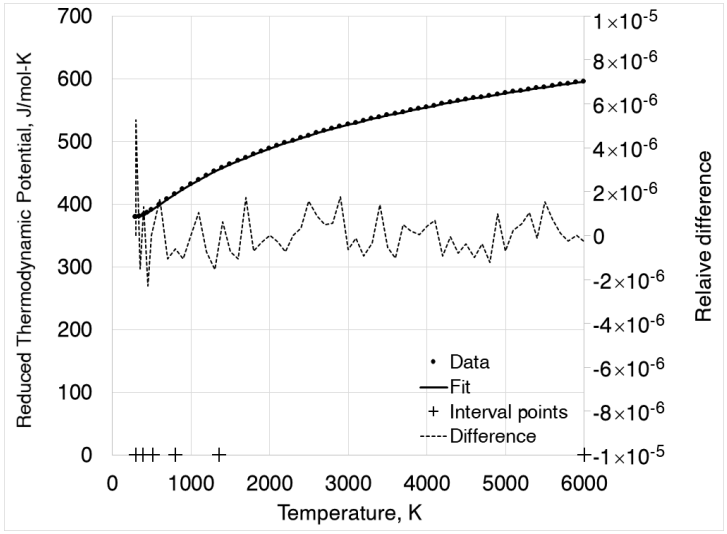
Tetrabromosilane ($SiBr_4$) – Br4Si1(g)

MW = 347.702 g/mol – $\Delta_f H^o$(298.15 K) = -415.471 kJ/mol – S^o(298.15 K) = 379.365 J/mol-K – p^o = 0.1 MPa



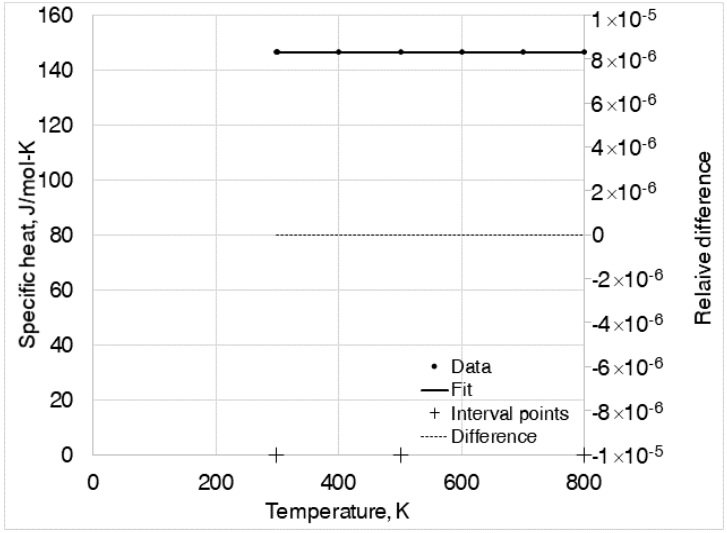


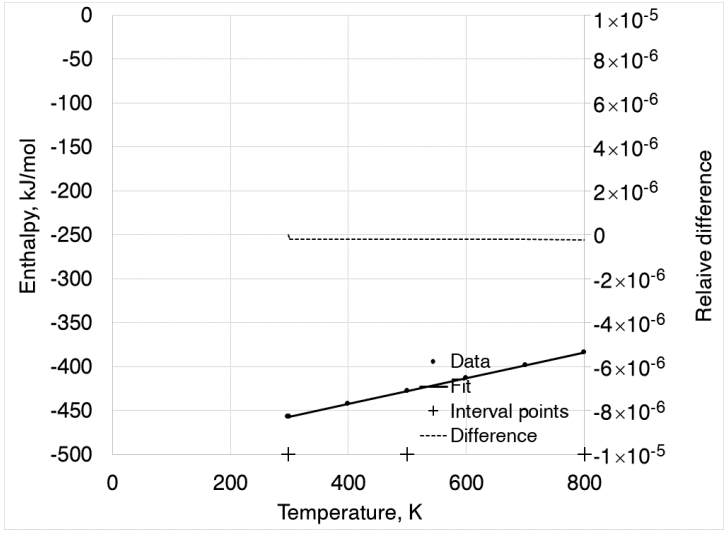


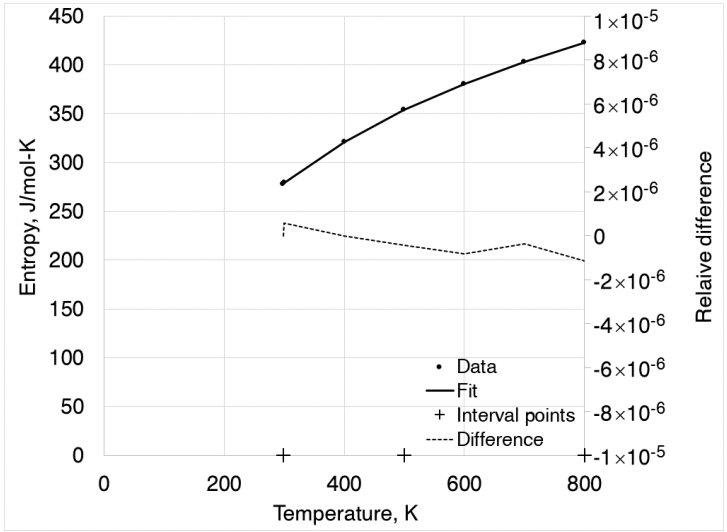


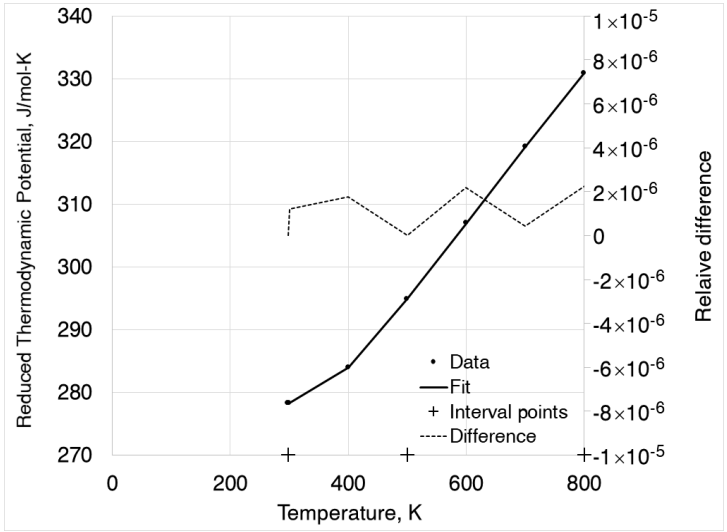
Tetrabromosilane ($SiBr_4$) – Br4Si1(l)

MW = 347.702 g/mol – $\Delta_f H^o$(298.15 K) = -457.311 kJ/mol – S^o(298.15 K) = 278.236 J/mol-K – p^o = 0.1 MPa



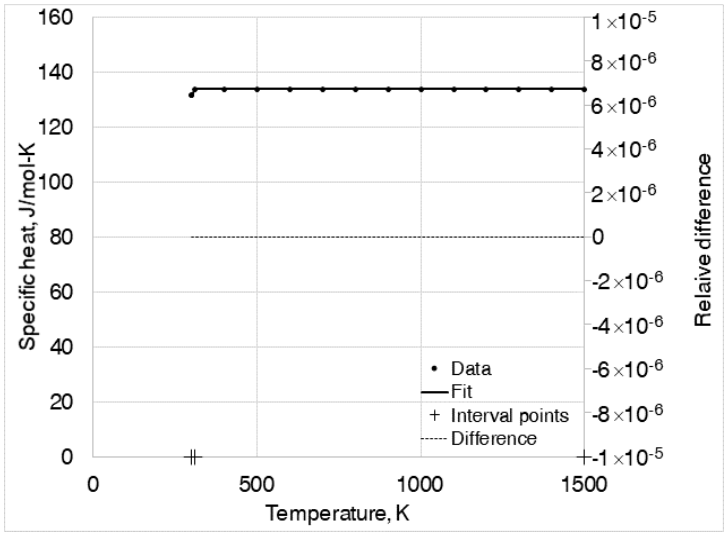


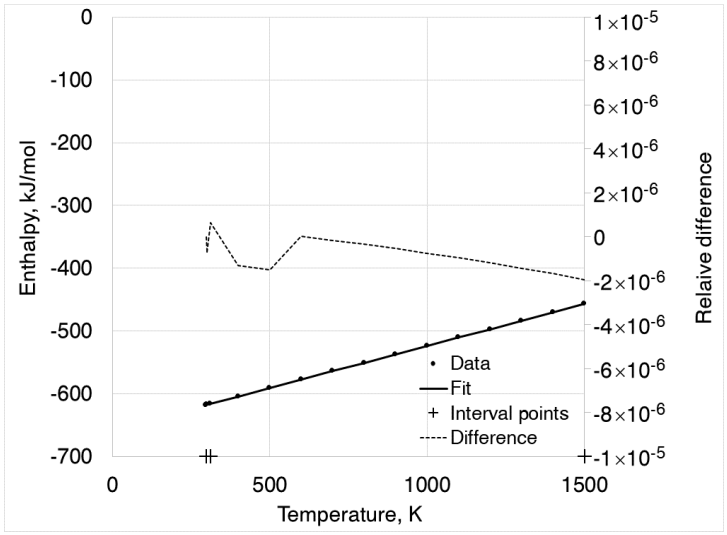


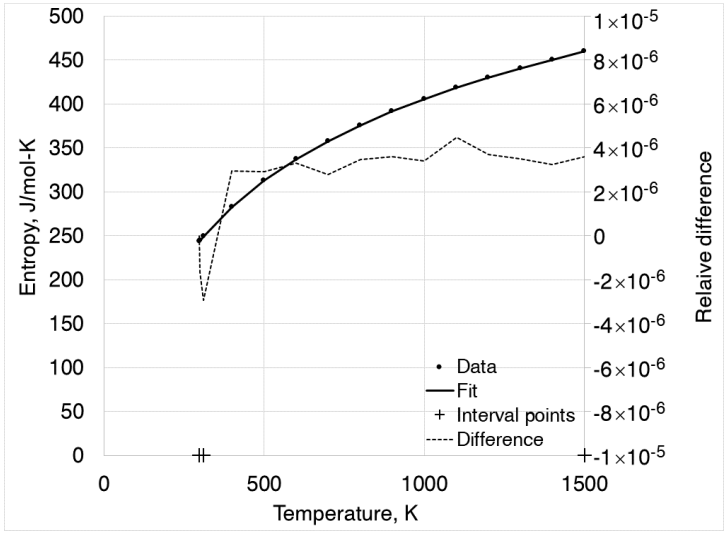


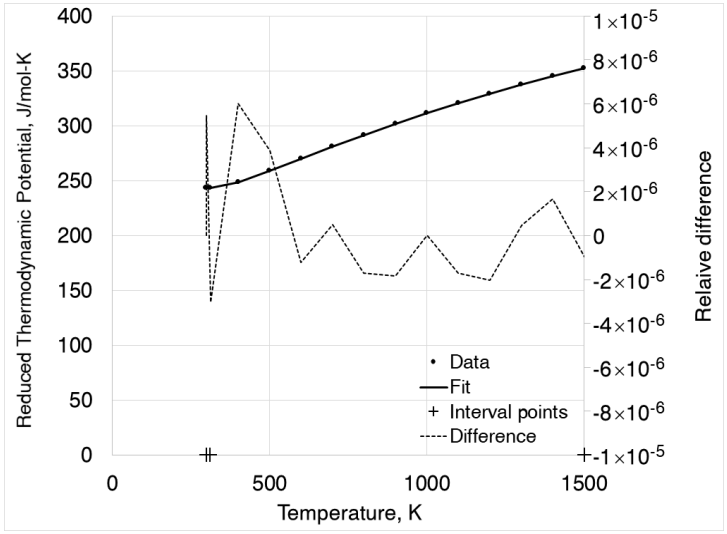
Titanium Bromide ($TiBr_4$) – Br4Ti1(cr)

MW = 367.5 g/mol – $\Delta_f H^o$(298.15 K) = -617.977 kJ/mol – S^o(298.15 K) = 243.622 J/mol-K – p^o = 0.1 MPa



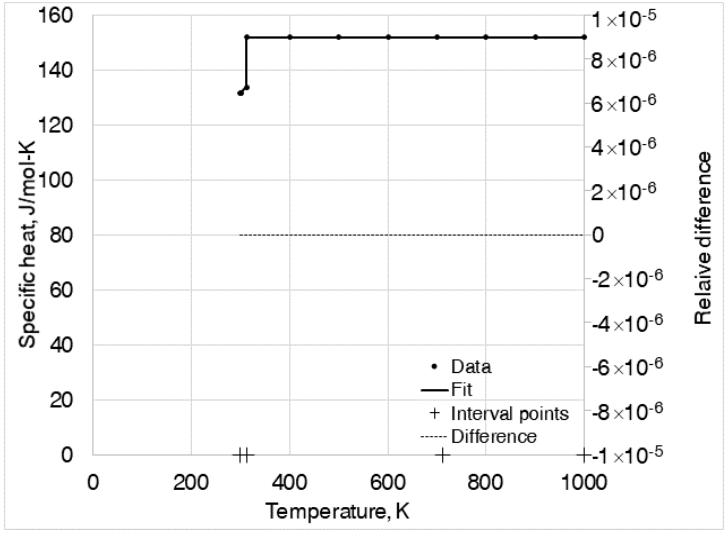


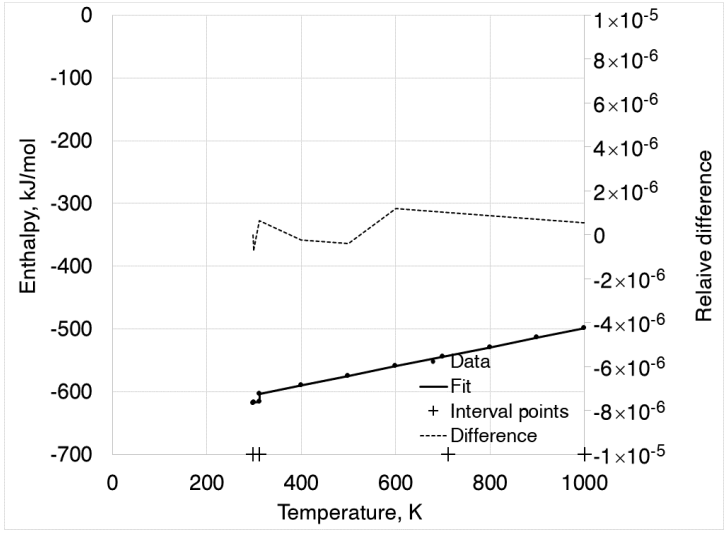


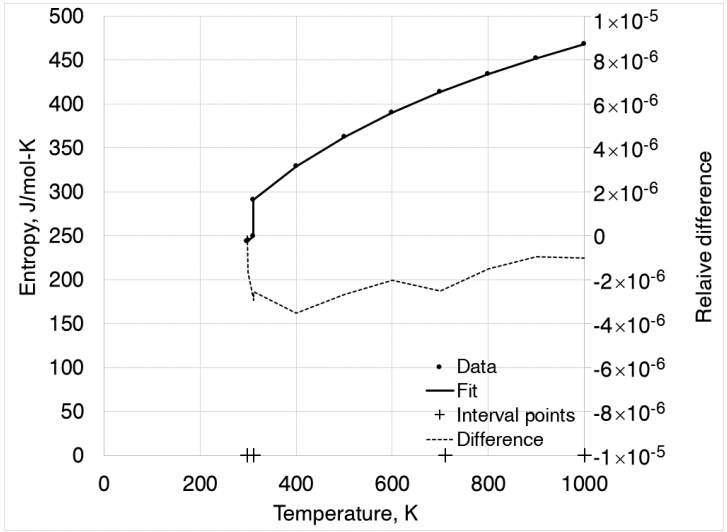


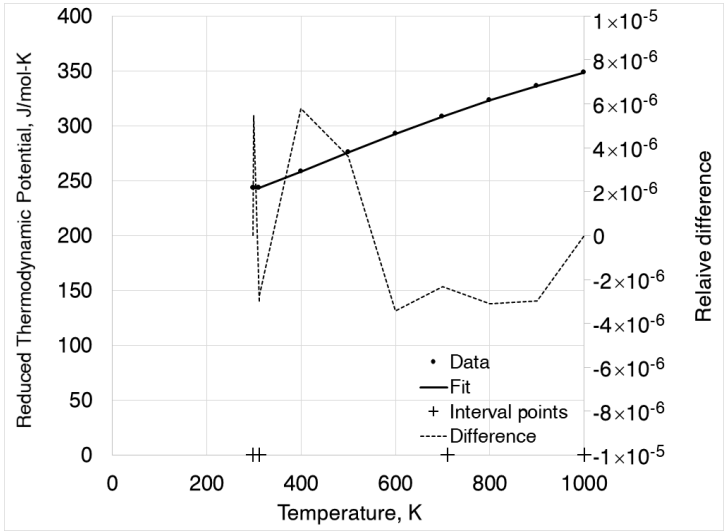
Titanium Bromide ($TiBr_4$) – Br4Ti1(cr,l)

MW = 367.5 g/mol – $\Delta_f H^o$(298.15 K) = -617.977 kJ/mol – S^o(298.15 K) = 243.622 J/mol-K – p^o = 0.1 MPa



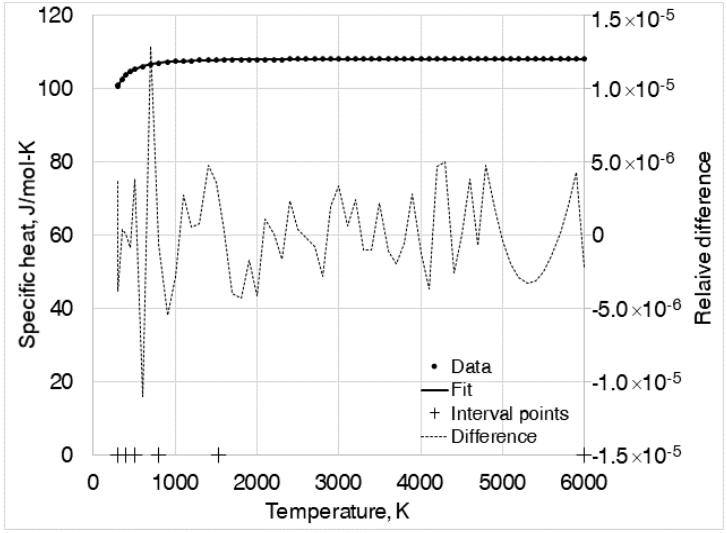


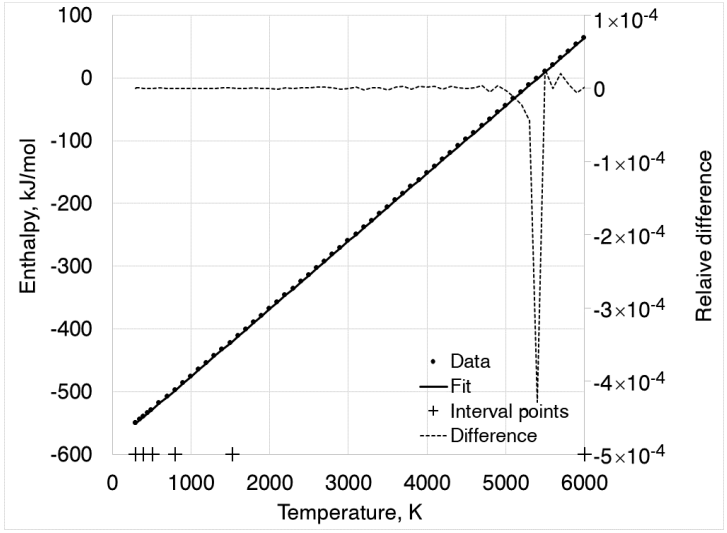


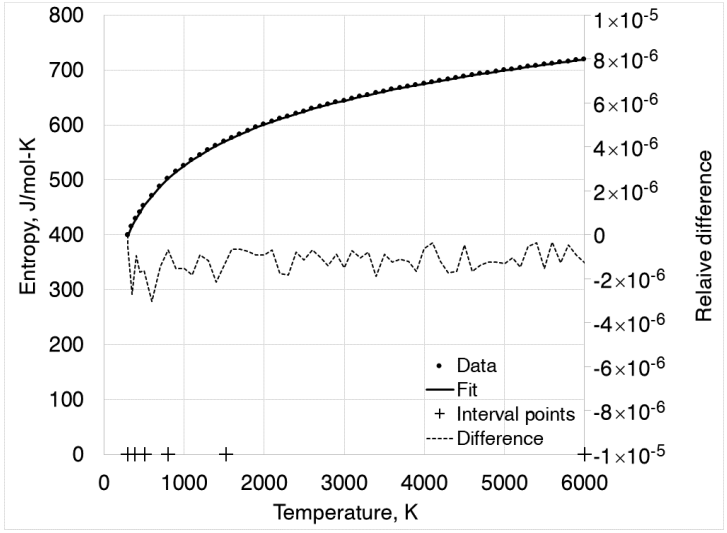


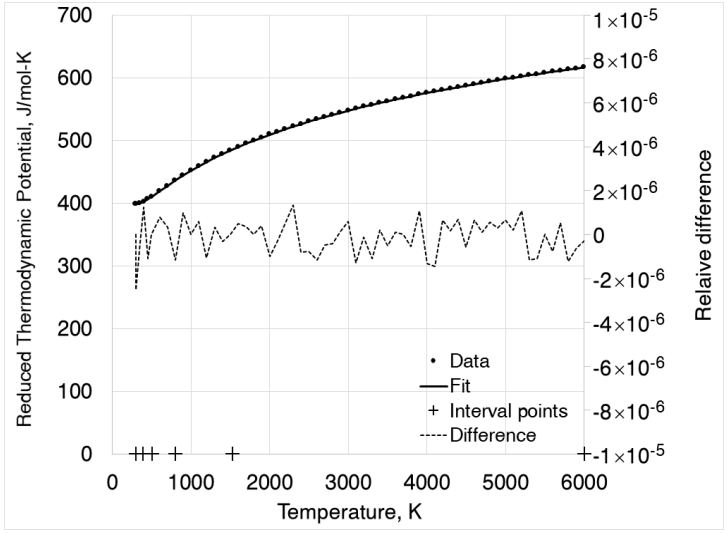
Titanium Bromide ($TiBr_4$) – Br4Ti1(g)

MW = 367.5 g/mol – $\Delta_f H^o$(298.15 K) = -550.196 kJ/mol – S^o(298.15 K) = 398.601 J/mol-K – p^o = 0.1 MPa



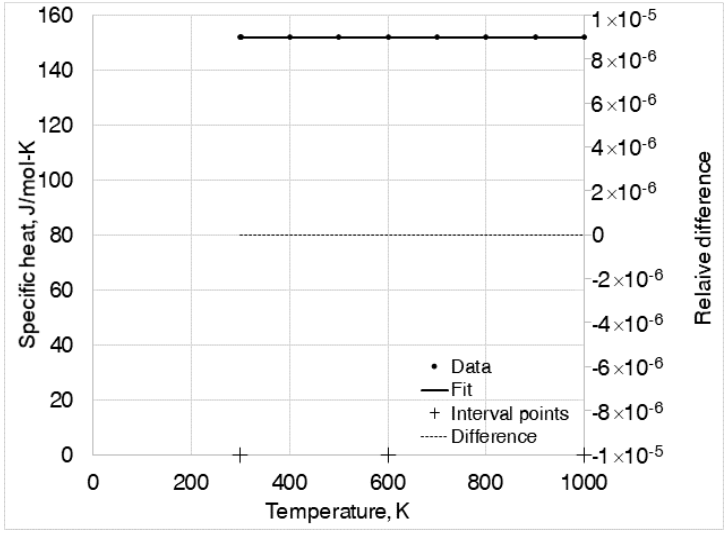


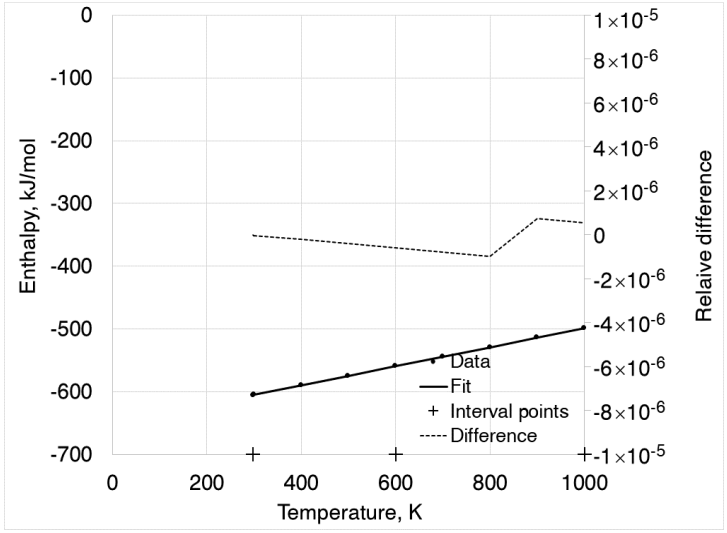


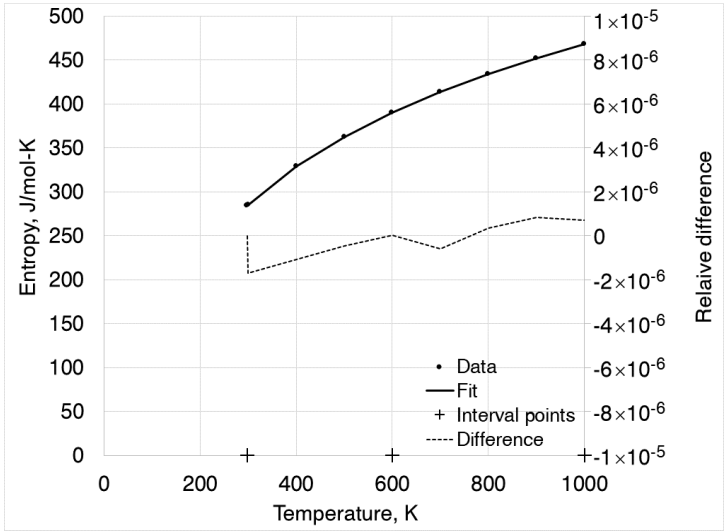


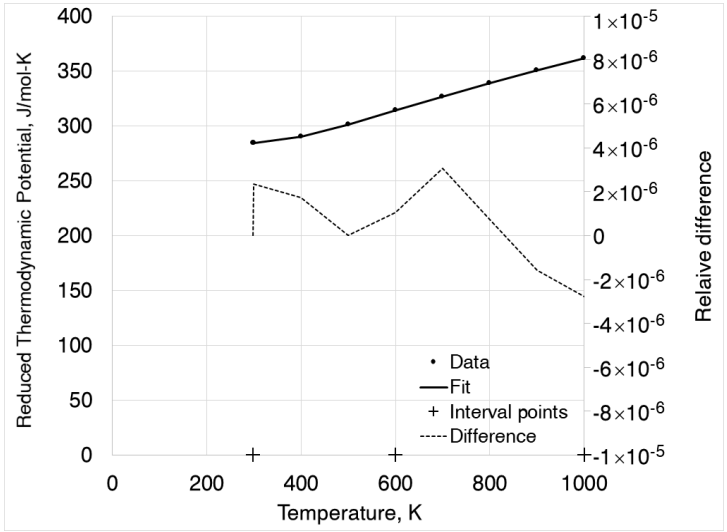
Titanium Bromide ($TiBr_4$) – Br4Ti1(l)

MW = 367.5 g/mol – $\Delta_f H^o$(298.15 K) = -605.345 kJ/mol – S^o(298.15 K) = 284.168 J/mol-K – p^o = 0.1 MPa



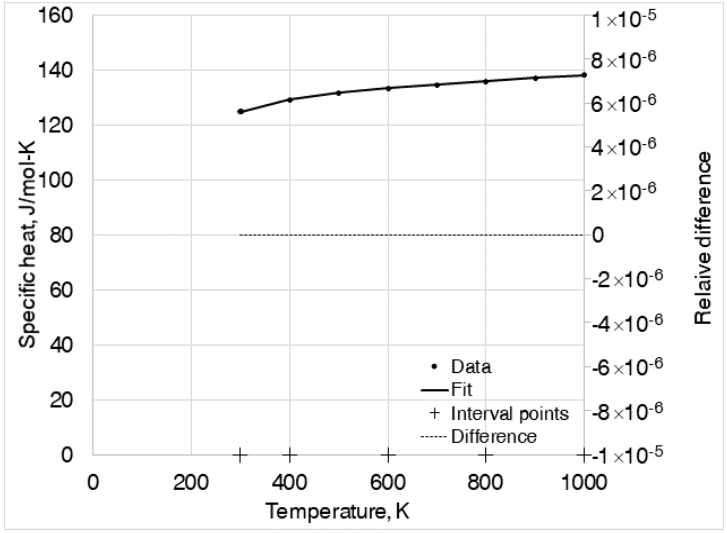


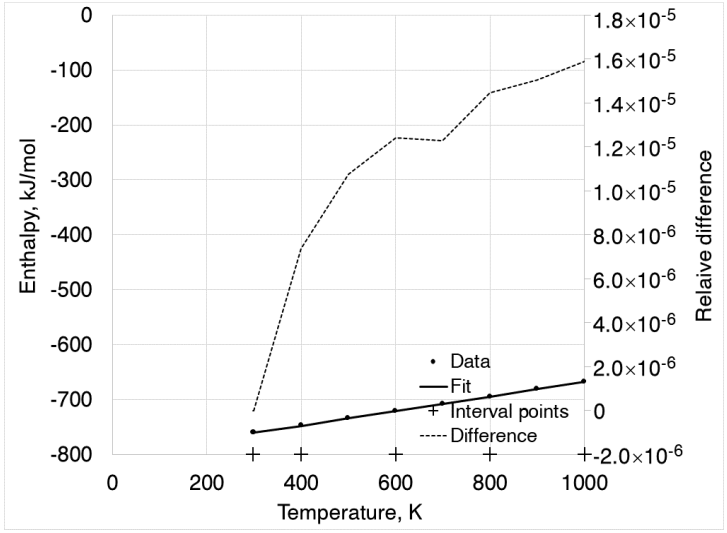


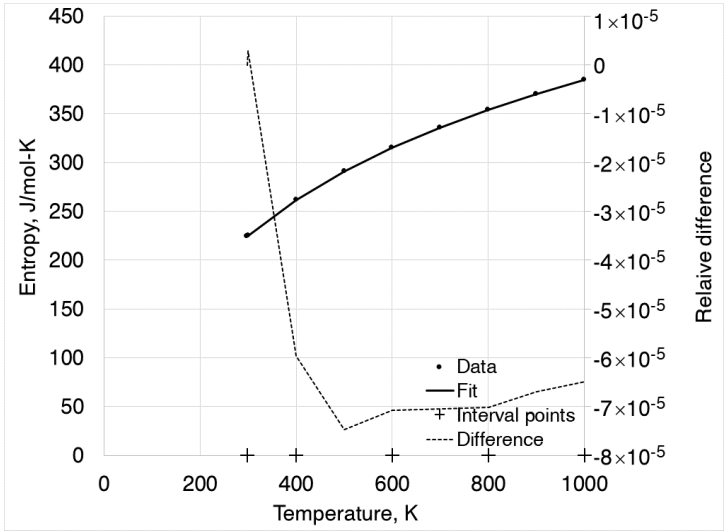


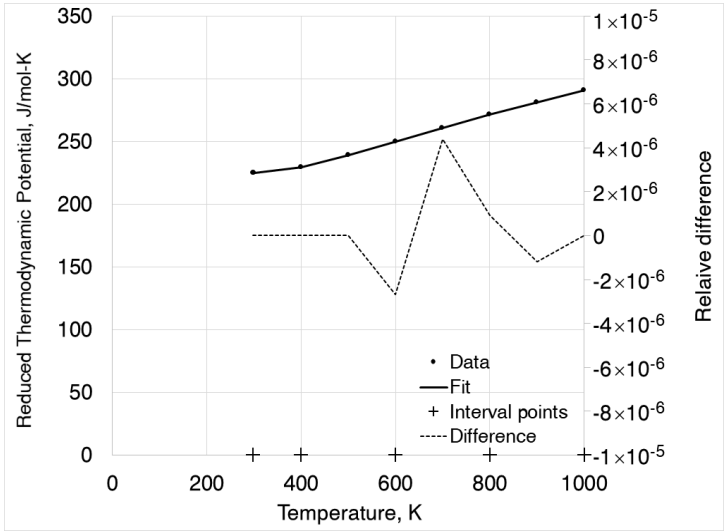
Zirconium Bromide ($ZrBr_4$) – Br4Zr1(cr)

MW = 410.84 g/mol – $\Delta_f H^o$(298.15 K) = -760.651 kJ/mol – S^o(298.15 K) = 224.689 J/mol-K – p^o = 0.1 MPa



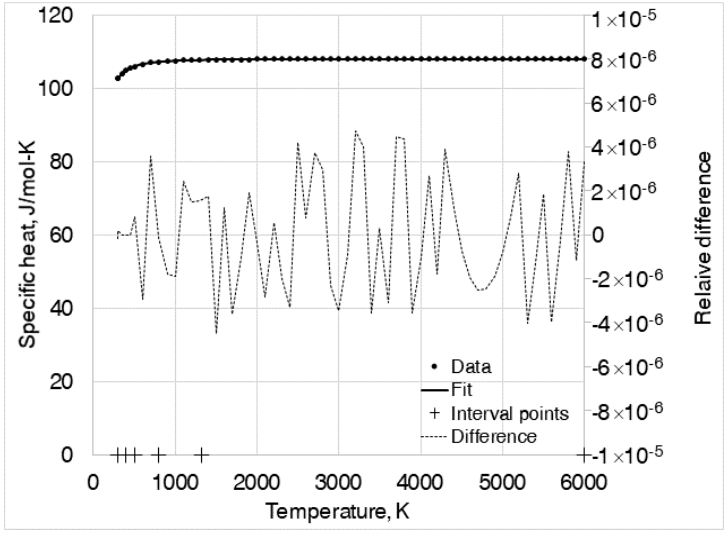


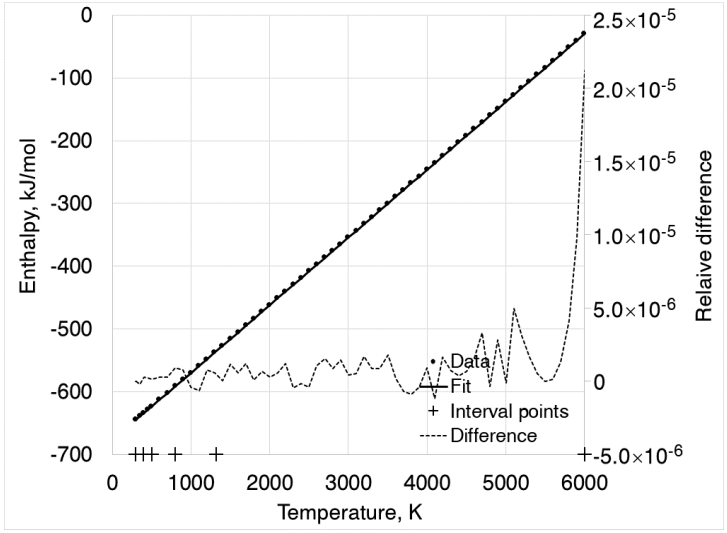


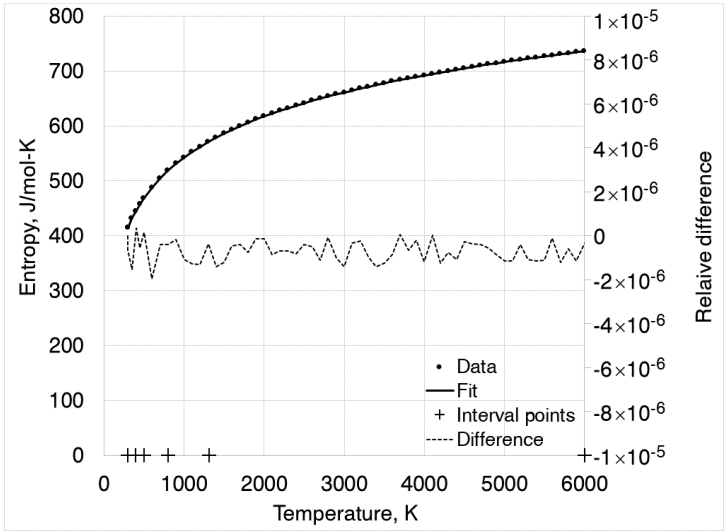


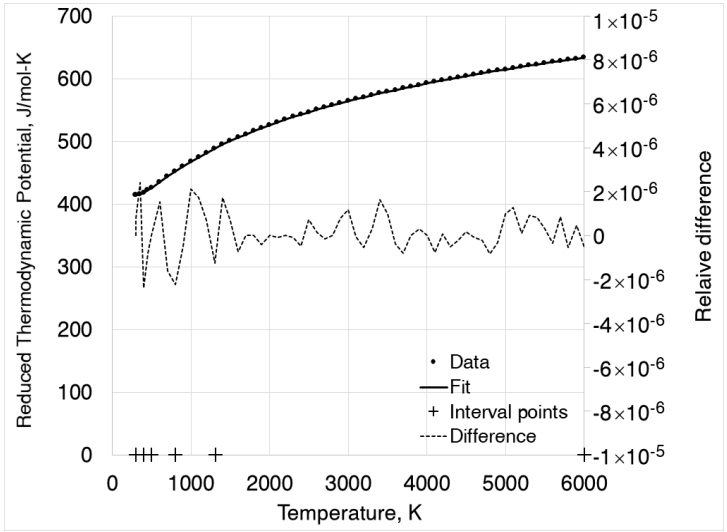
Zirconium Bromide ($ZrBr_4$) – Br4Zr1(g)

MW = 410.84 g/mol – $\Delta_f H^o$(298.15 K) = -644.754 kJ/mol – S^o(298.15 K) = 414.487 J/mol-K – p^o = 0.1 MPa



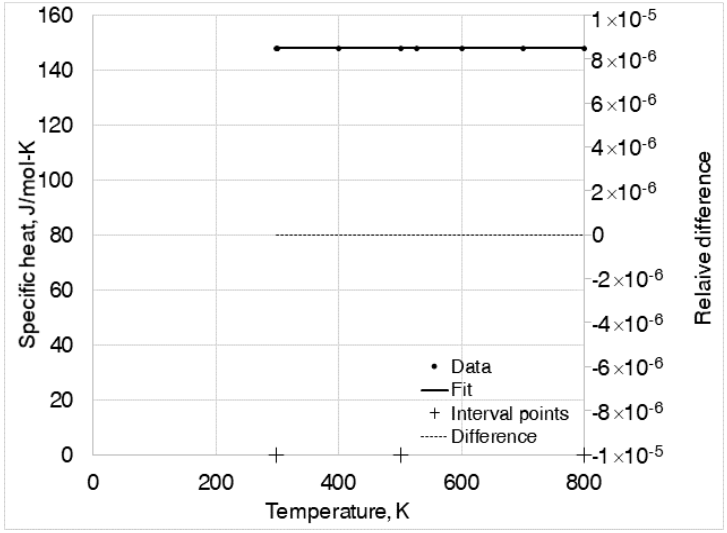


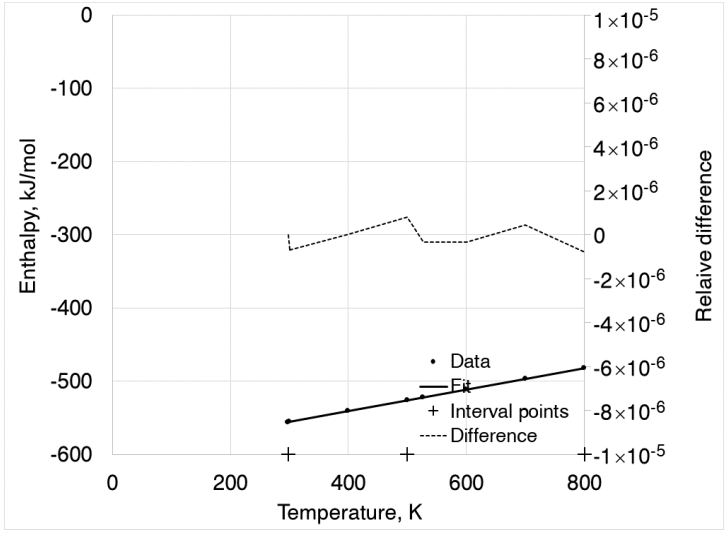


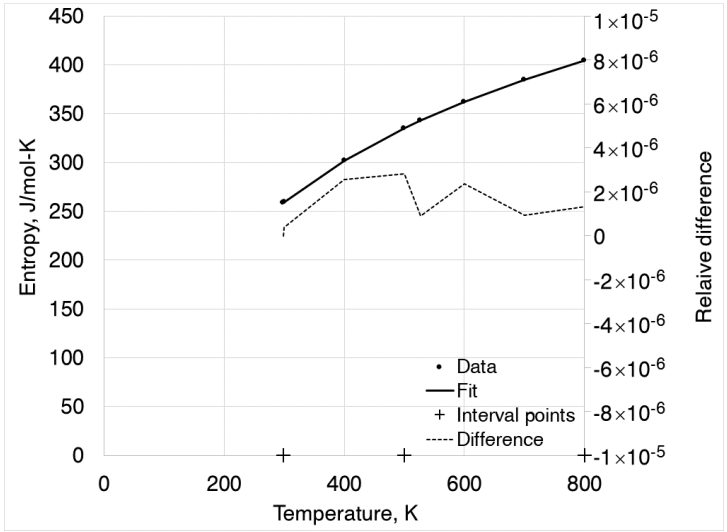


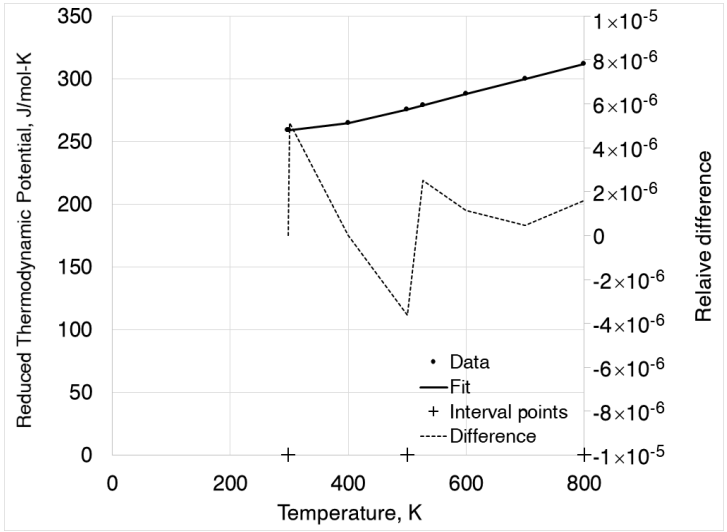
Niobium Bromide ($NbBr_5$) – Br5Nb1(cr)

MW = 492.426 g/mol – $\Delta_f H^o$(298.15 K) = -556.054 kJ/mol – S^o(298.15 K) = 258.78 J/mol-K – p^o = 0.1 MPa



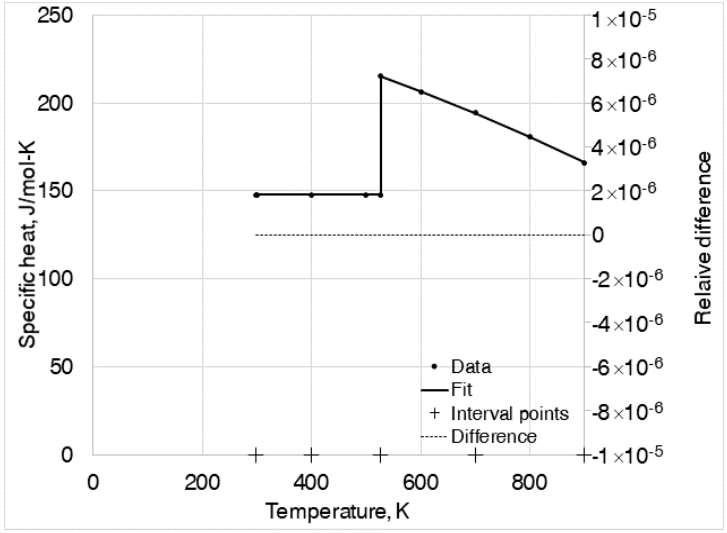


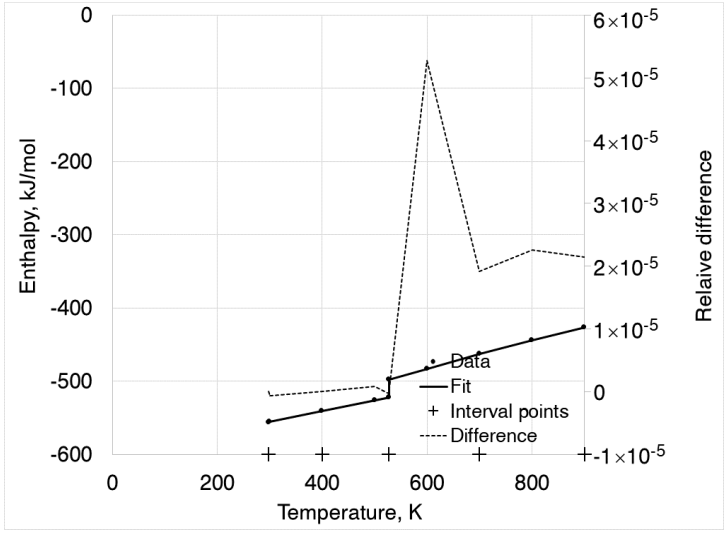


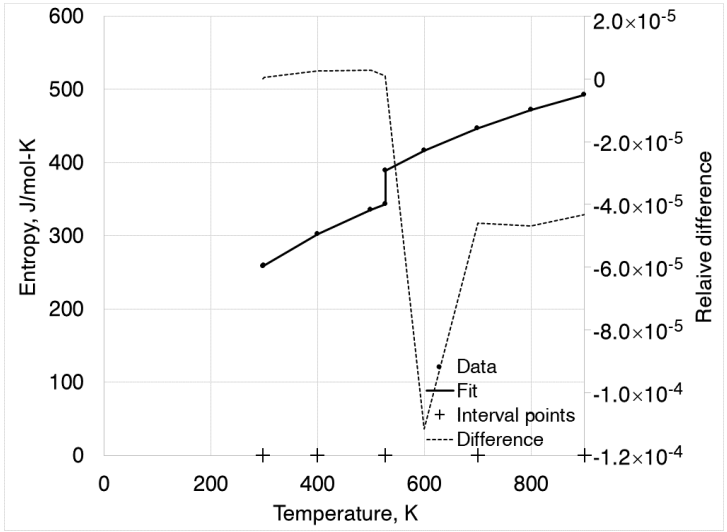


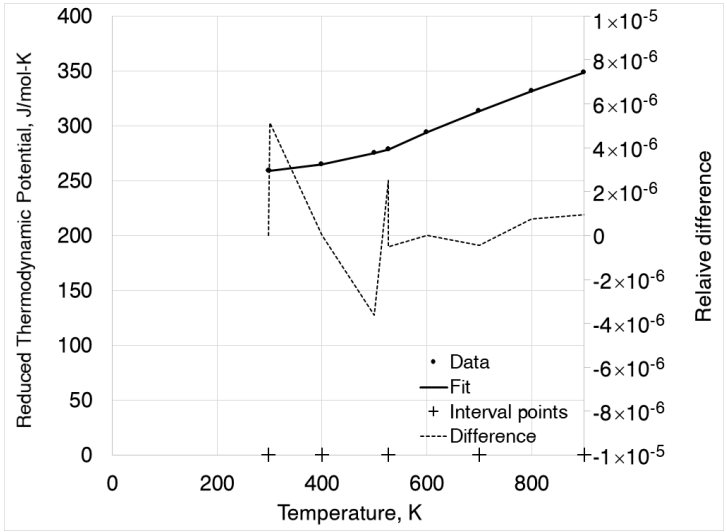
Niobium Bromide ($NbBr_5$) – Br5Nb1(cr,l)

MW = 492.426 g/mol – $\Delta_f H^o$(298.15 K) = -556.054 kJ/mol – S^o(298.15 K) = 258.78 J/mol-K – p^o = 0.1 MPa



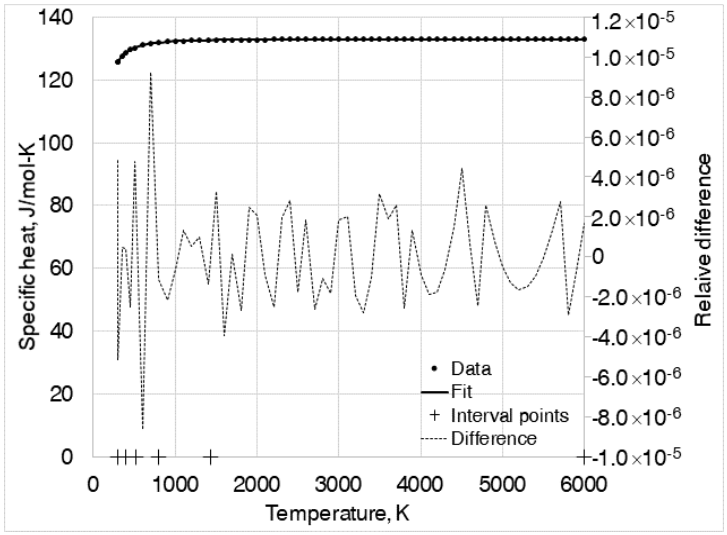


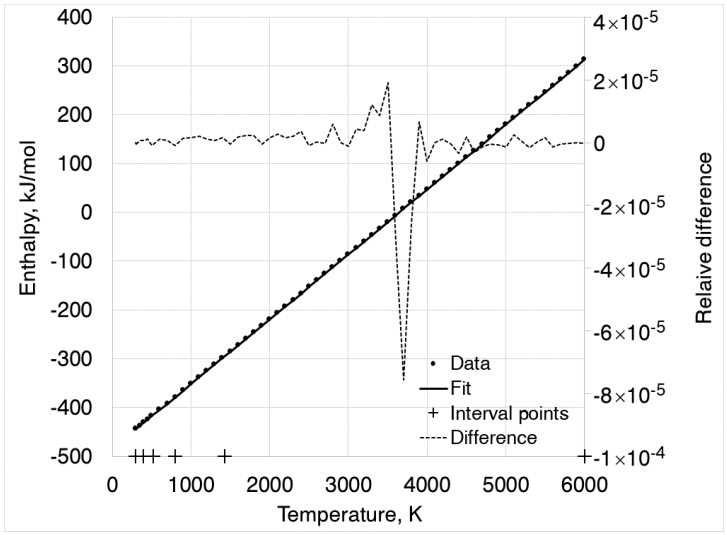


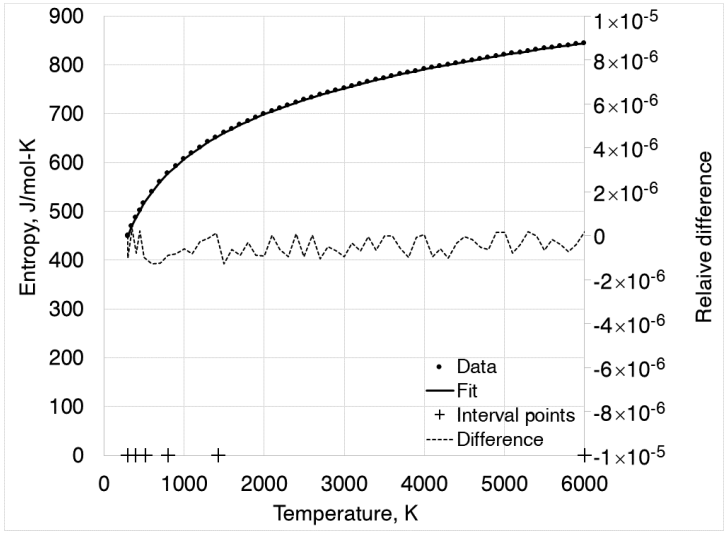


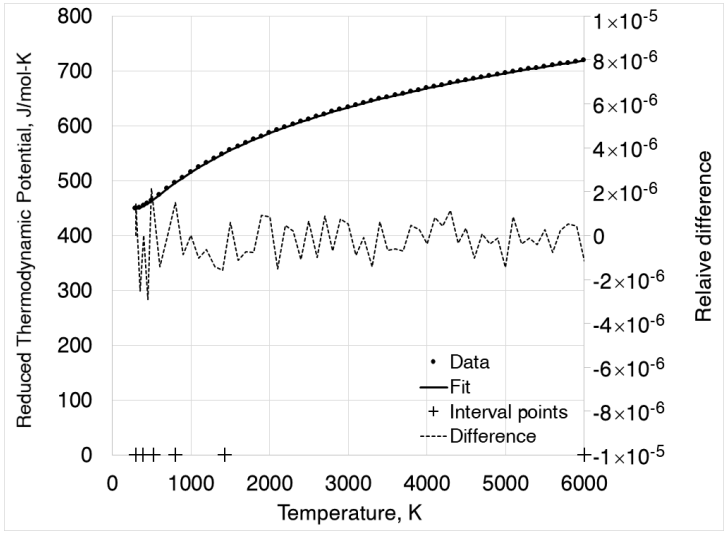
Niobium Bromide ($NbBr_5$) – Br5Nb1(g)

MW = 492.426 g/mol – $\Delta_f H^o$(298.15 K) = -443.588 kJ/mol – S^o(298.15 K) = 449.259 J/mol-K – p^o = 0.1 MPa



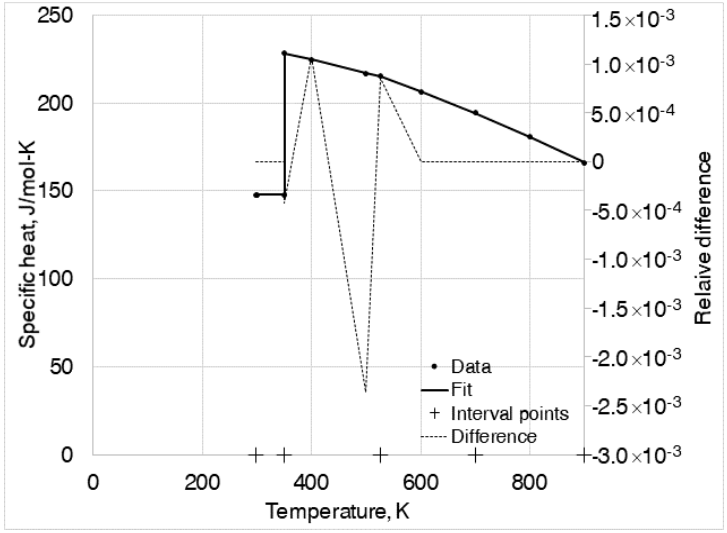


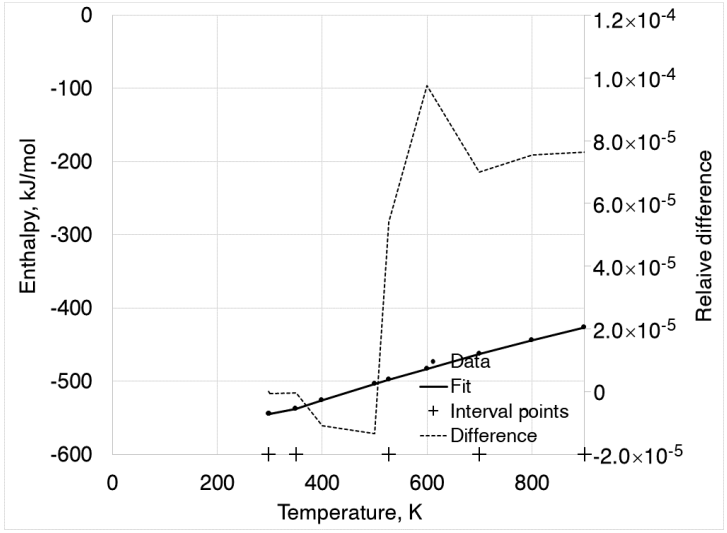


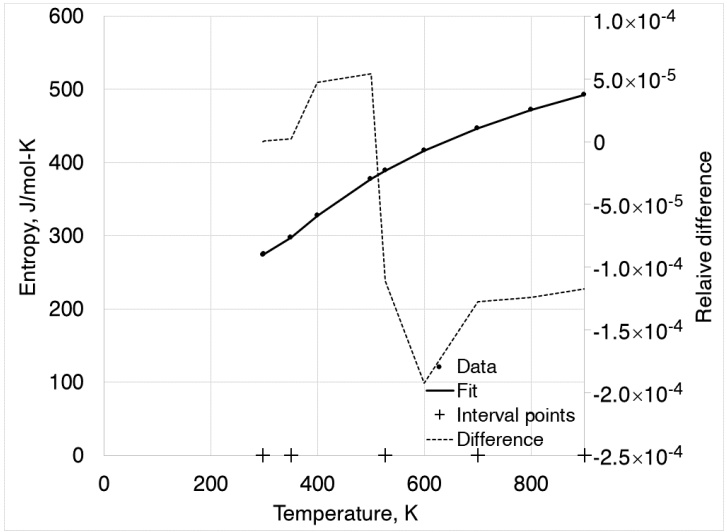


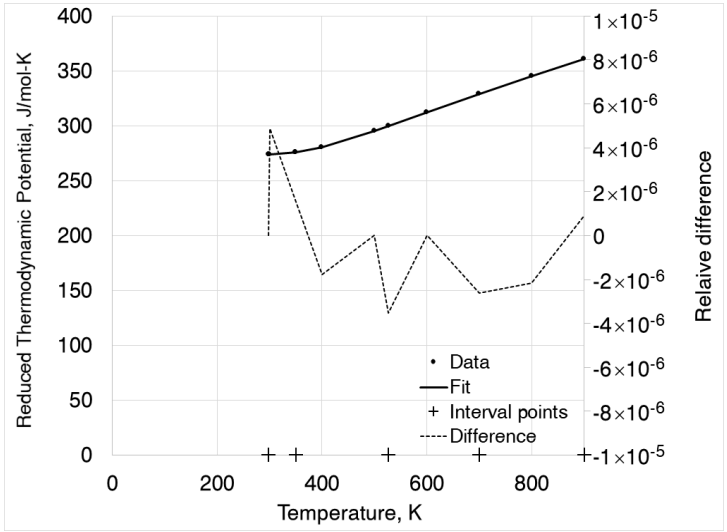
Niobium Bromide ($NbBr_5$) – Br5Nb1(l)

MW = 492.426 g/mol – $\Delta_f H^o$(298.15 K) = -545.074 kJ/mol – S^o(298.15 K) = 274.005 J/mol-K – p^o = 0.1 MPa



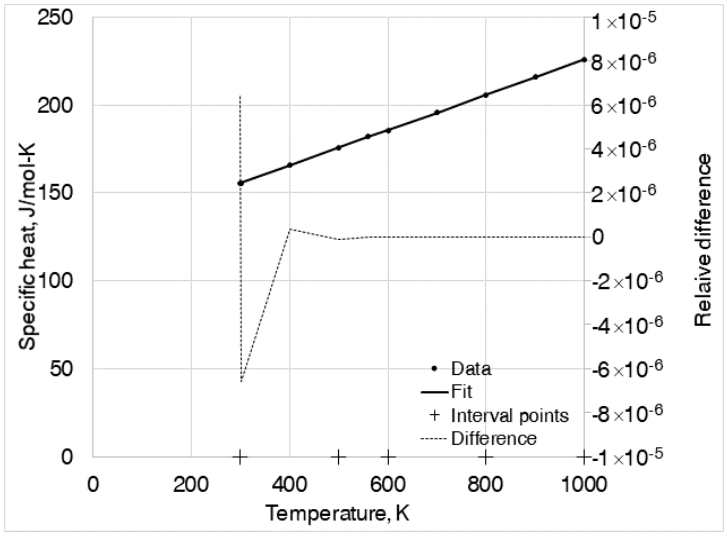


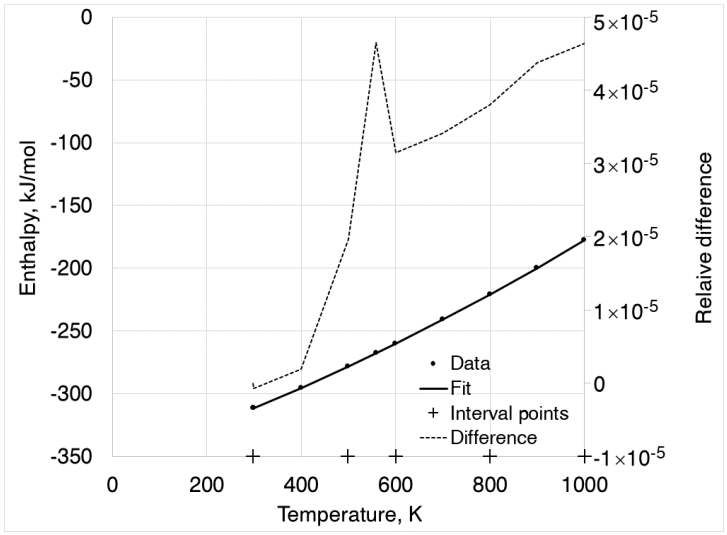


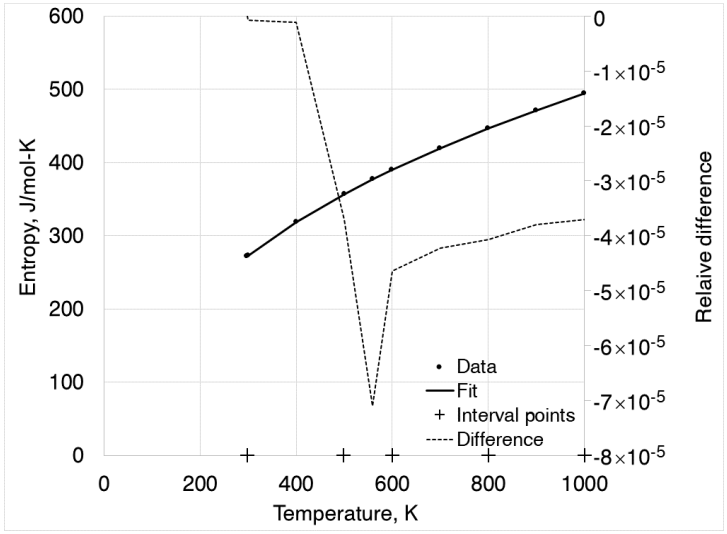


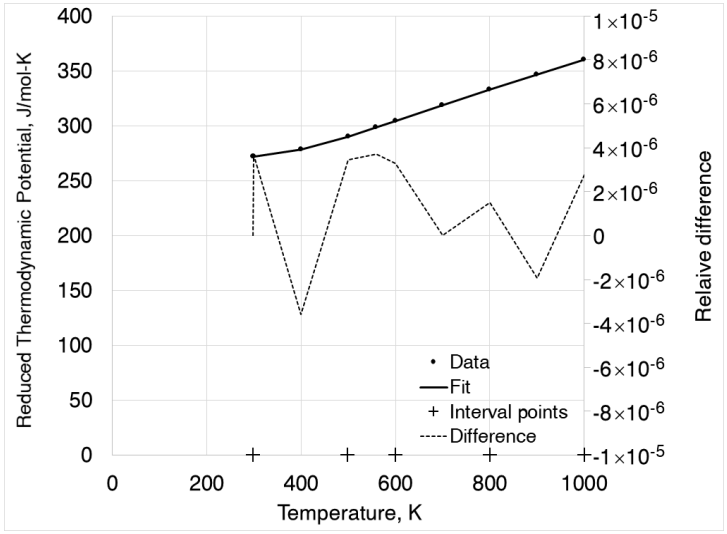
Tungsten Bromide (WBr_5) – Br5W1(cr)

MW = 583.37 g/mol – $\Delta_f H^o$(298.15 K) = -311.708 kJ/mol – S^o(298.15 K) = 271.96 J/mol-K – p^o = 0.1 MPa



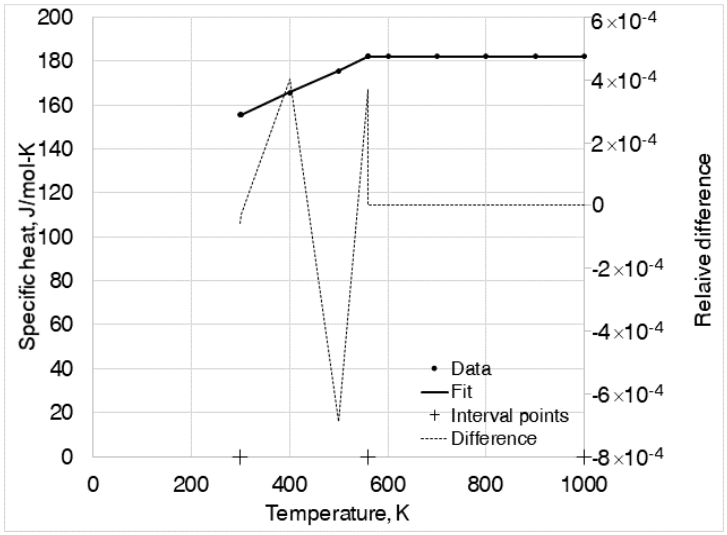


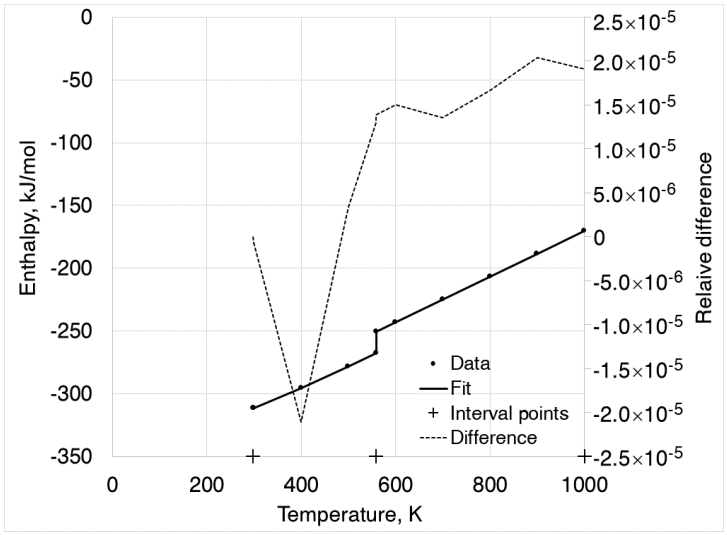


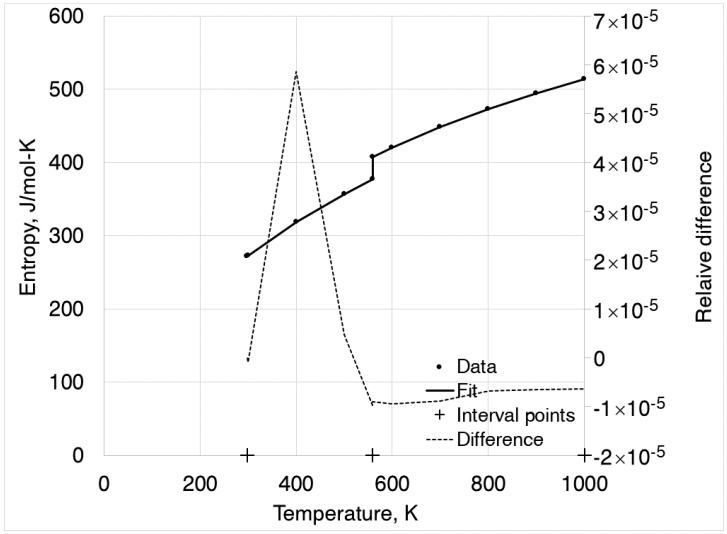


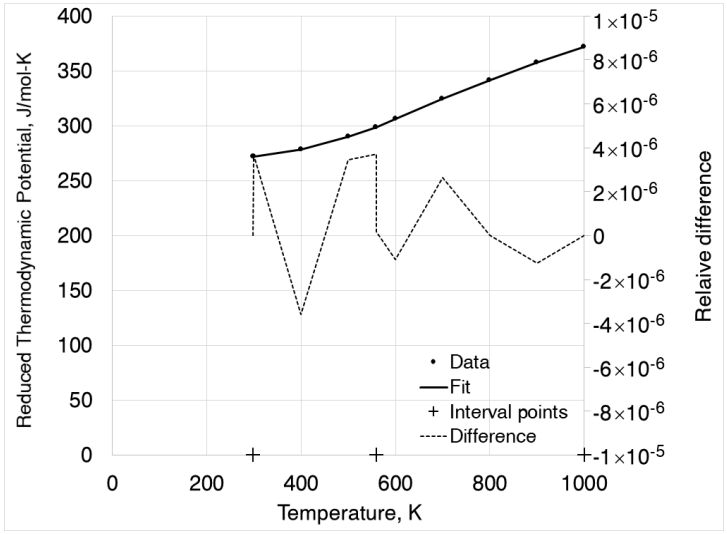
Tungsten Bromide (WBr_5) – Br5W1(cr,l)

MW = 583.37 g/mol – $\Delta_f H^o$(298.15 K) = -311.708 kJ/mol – S^o(298.15 K) = 271.96 J/mol-K – p^o = 0.1 MPa



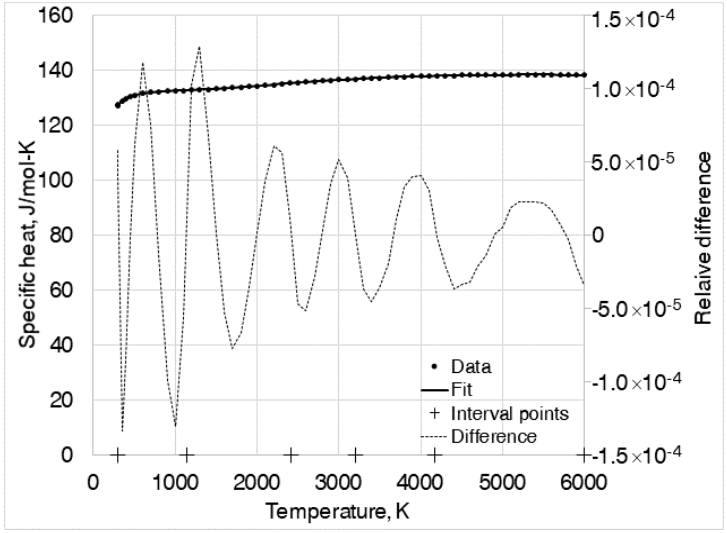


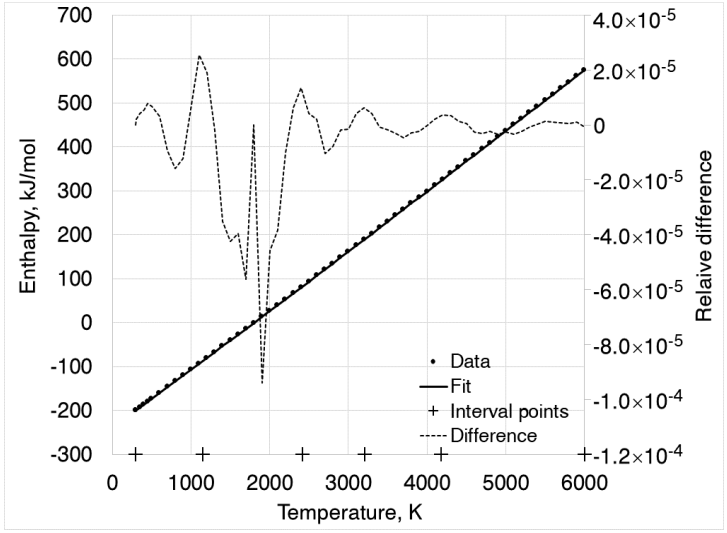


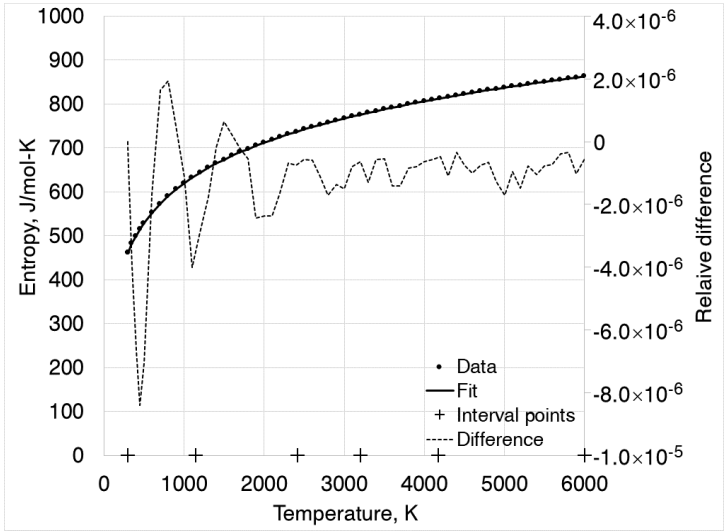


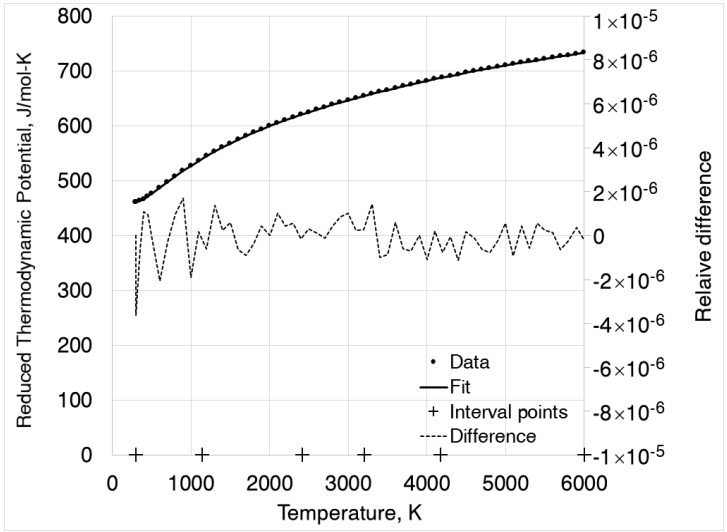
Tungsten Bromide (WBr_5) – Br5W1(g)

MW = 583.37 g/mol – $\Delta_f H^o$(298.15 K) = -199.158 kJ/mol – S^o(298.15 K) = 461.411 J/mol-K – p^o = 0.1 MPa



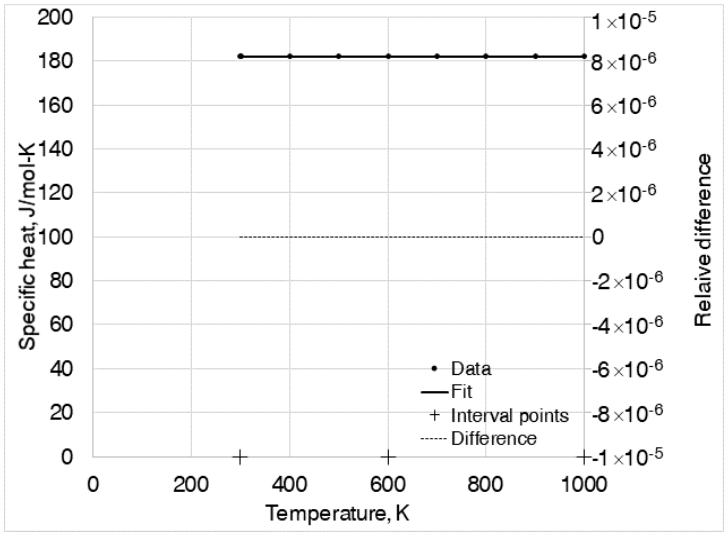


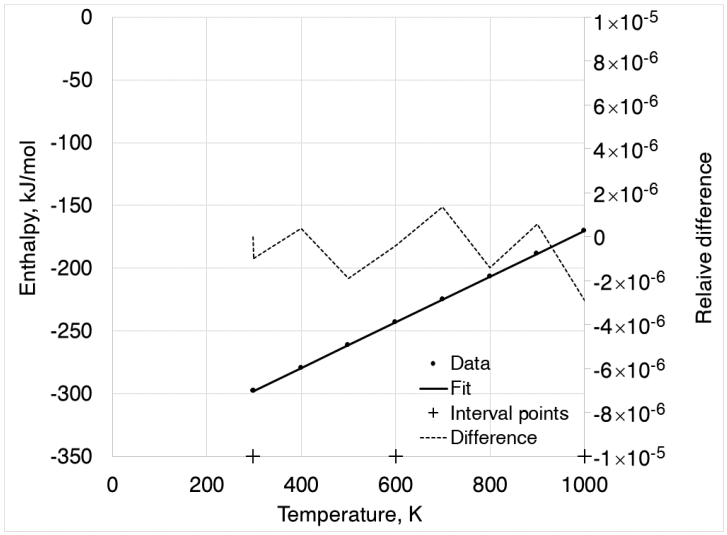


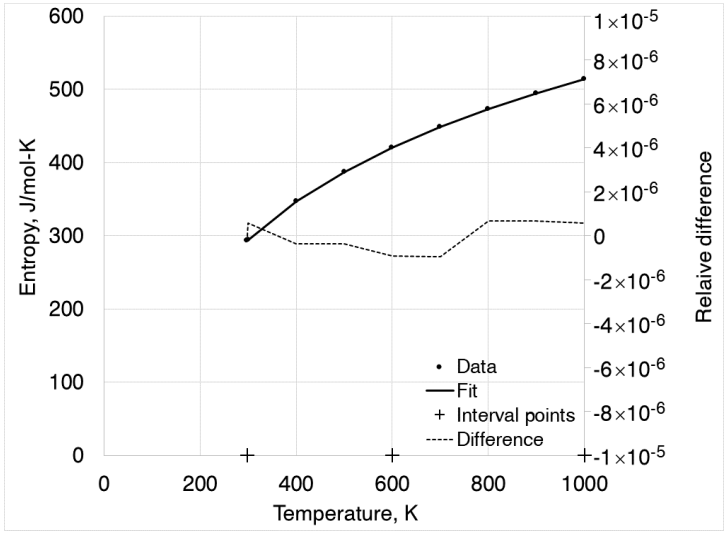


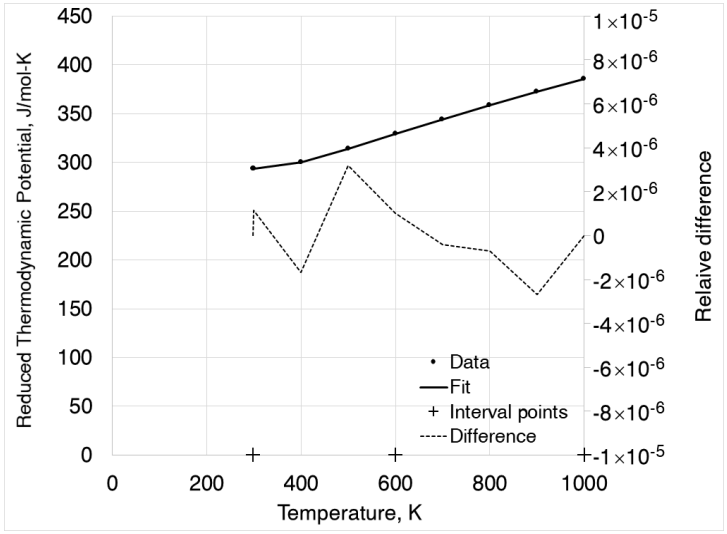
Tungsten Bromide (WBr_5) – Br5W1(l)

MW = 583.37 g/mol – $\Delta_f H^o$(298.15 K) = -298.059 kJ/mol – S^o(298.15 K) = 293.341 J/mol-K – p^o = 0.1 MPa



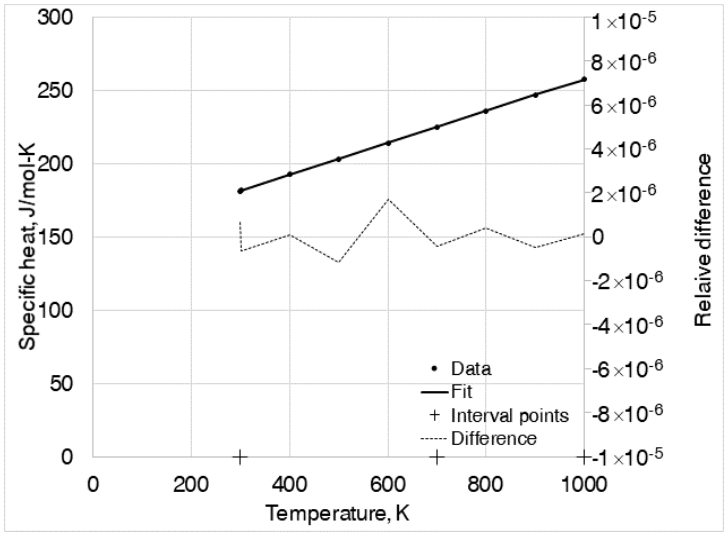


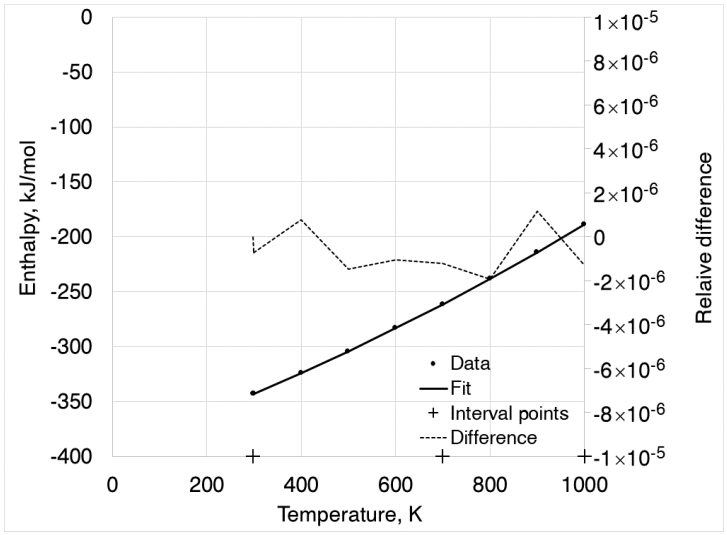


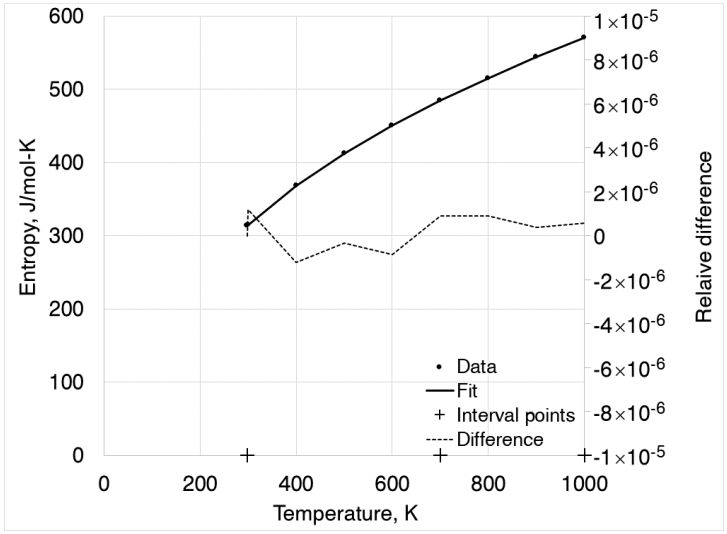


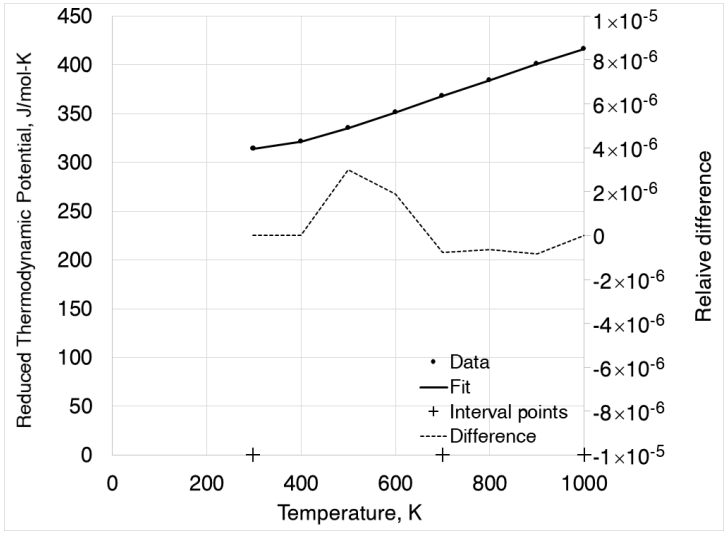
Tungsten Bromide (WBr_6) – Br6W1(cr)

MW = 663.27 g/mol – $\Delta_f H^o$(298.15 K) = -343.088 kJ/mol – S^o(298.15 K) = 313.8 J/mol-K – p^o = 0.1 MPa



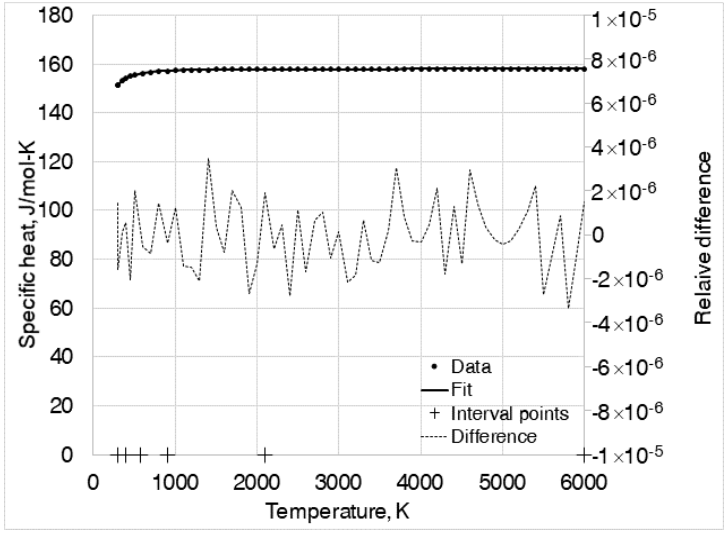


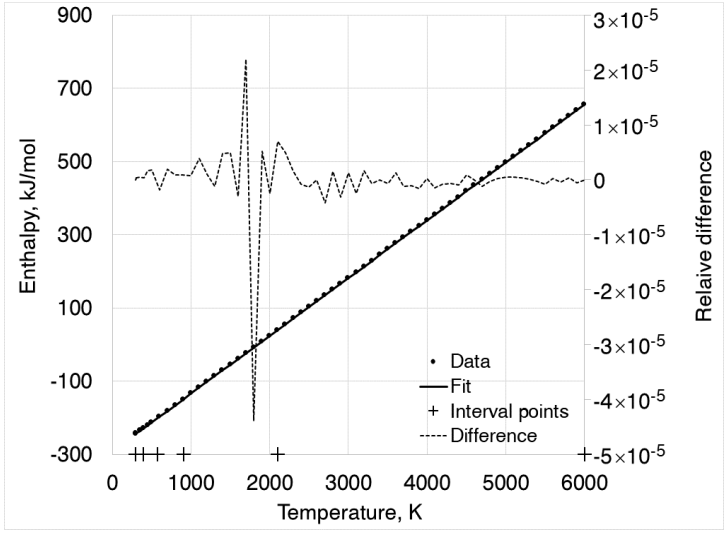


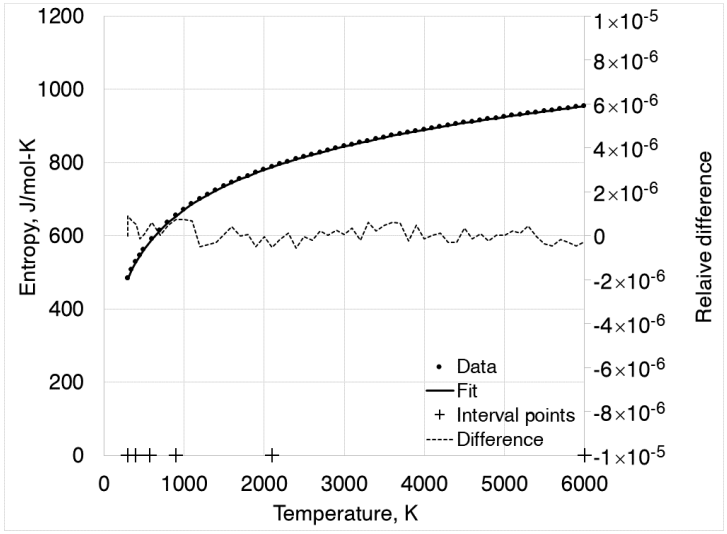


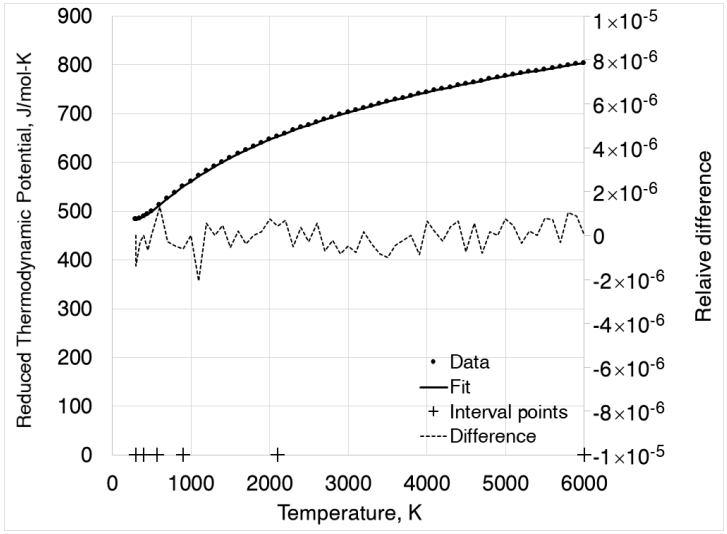
Tungsten Bromide (WBr_6) – Br6W1(g)

MW = 663.27 g/mol – $\Delta_f H^o$(298.15 K) = -243.09 kJ/mol – S^o(298.15 K) = 482.699 J/mol-K – p^o = 0.1 MPa



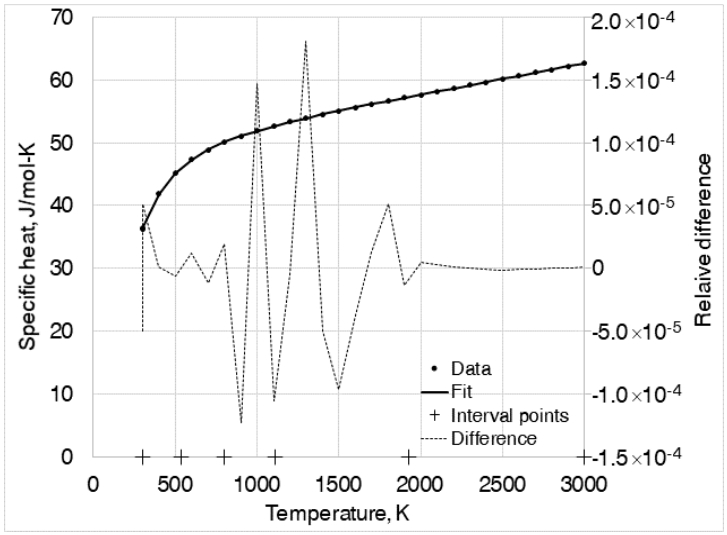


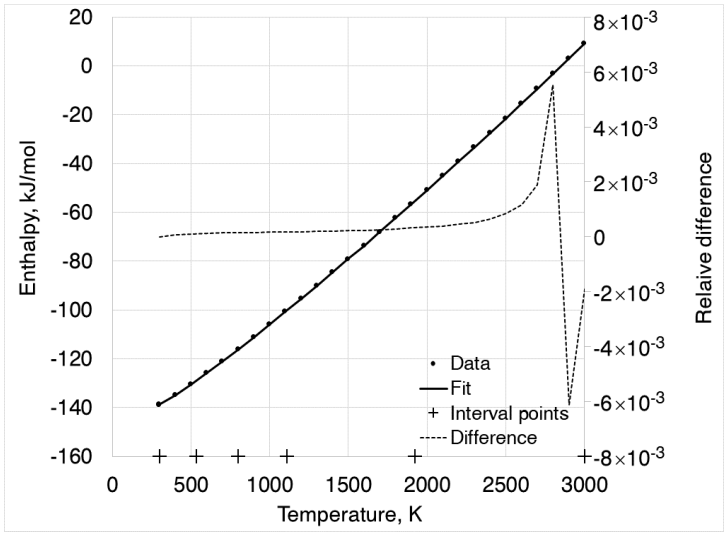


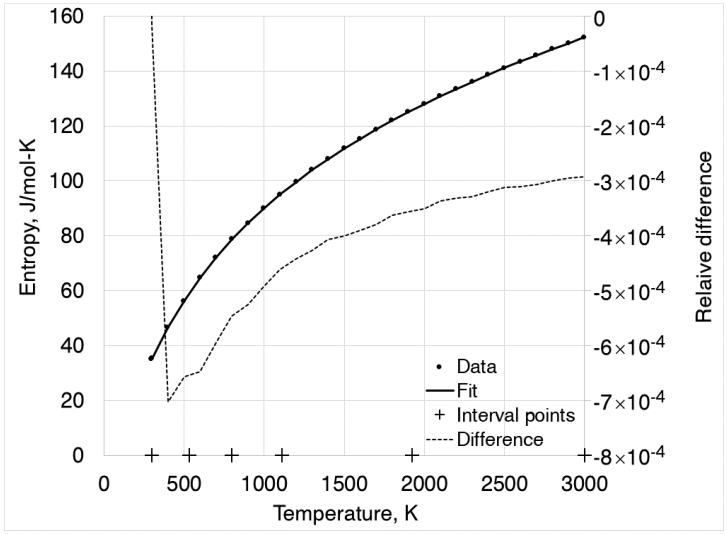


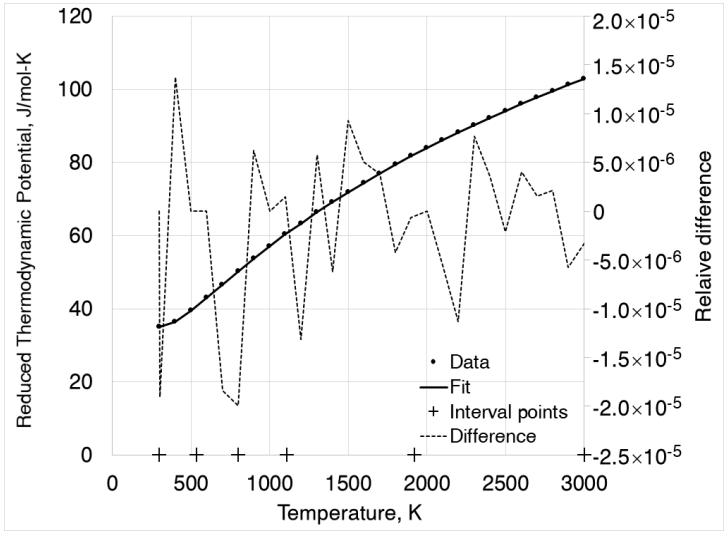
Niobium Carbide ($NbC_{0.98}$) – C0.98Nb1(cr)

MW = 104.677 g/mol – $\Delta_f H^o$(298.15 K) = -138.909 kJ/mol – S^o(298.15 K) = 34.966 J/mol-K – p^o = 0.1 MPa



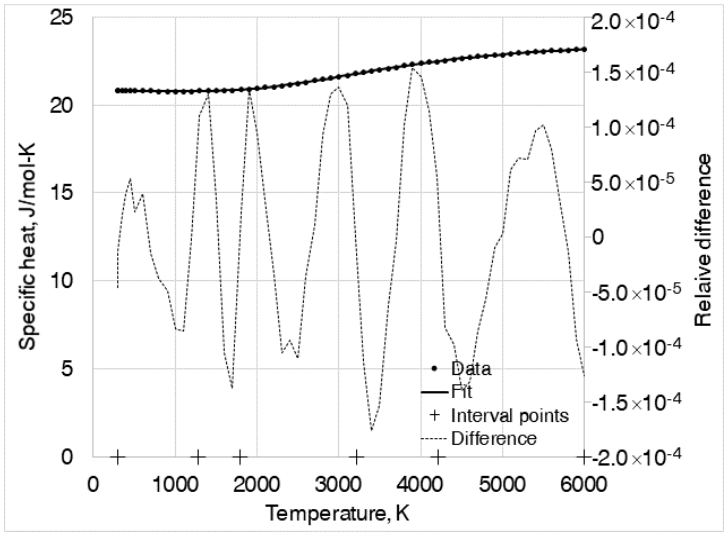


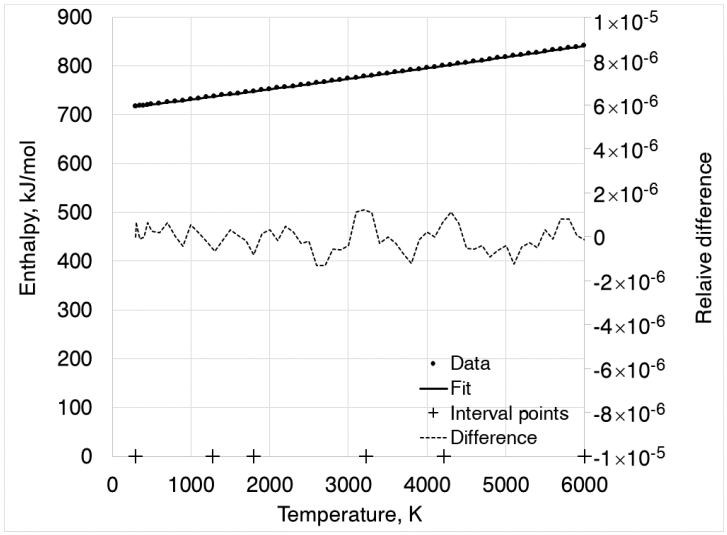


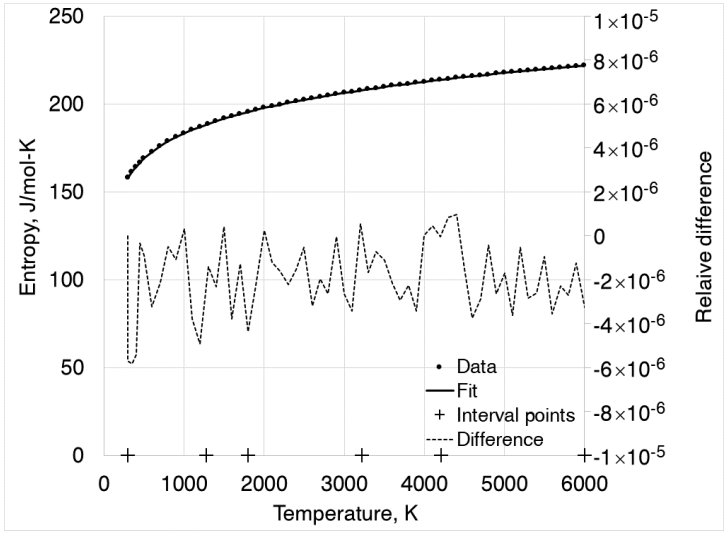


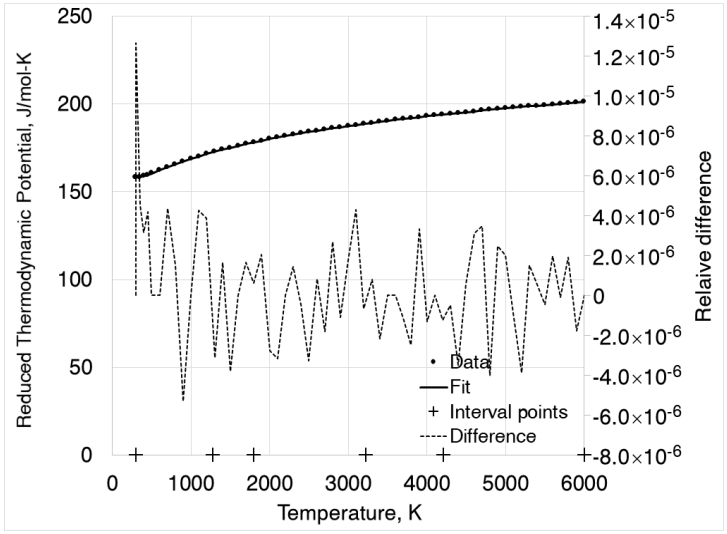
Carbon (C) – C1(g)

MW = 12.011 g/mol – $\Delta_f H^o$(298.15 K) = 716.67 kJ/mol – S^o(298.15 K) = 158.1 J/mol-K – p^o = 0.1 MPa



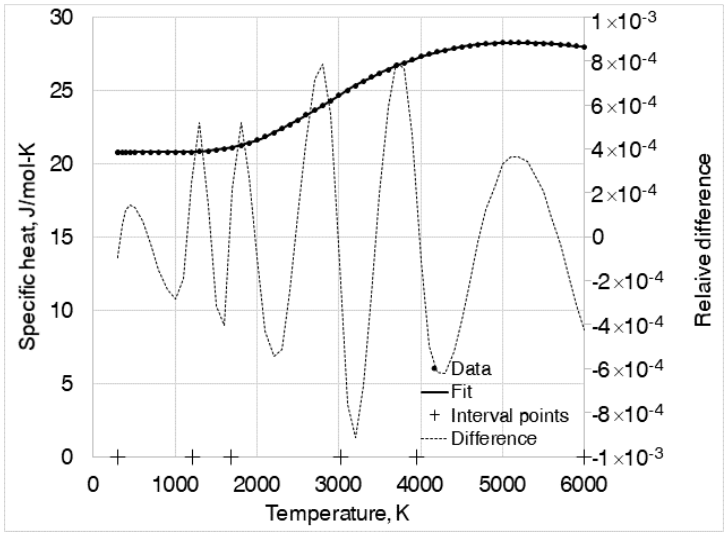


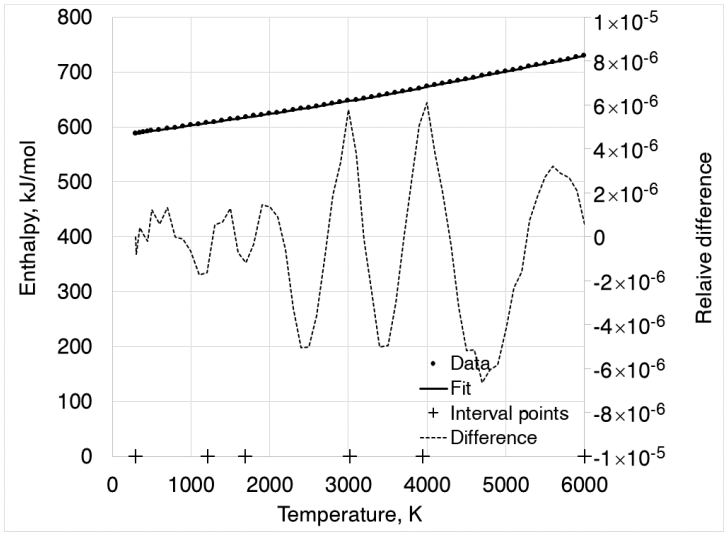


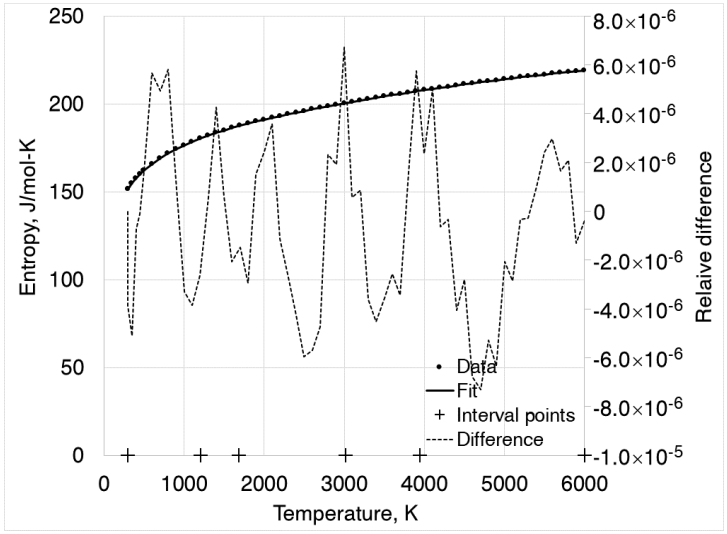


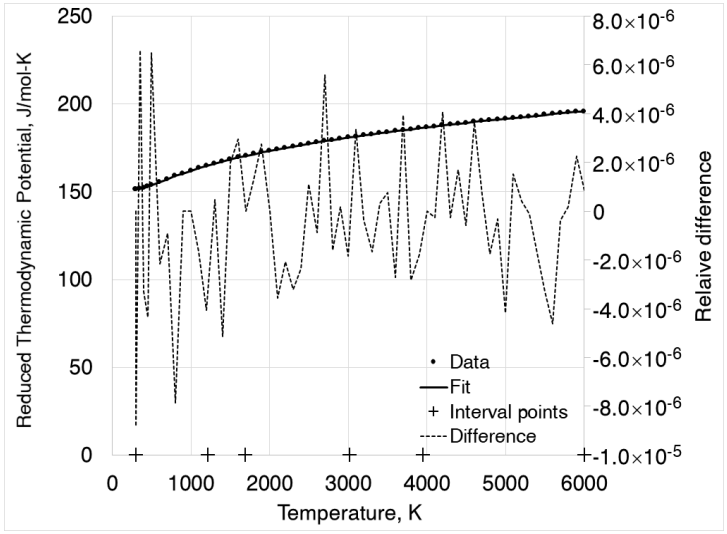
Carbon, Ion (C^-) – C1-(g)

MW = 12.012 g/mol – $\Delta_f H^o$(298.15 K) = 588.284 kJ/mol – S^o(298.15 K) = 151.384 J/mol-K – p^o = 0.1 MPa



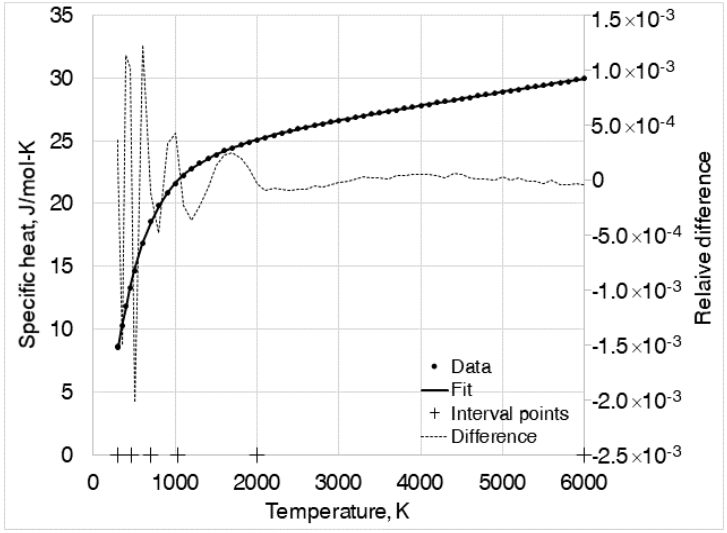


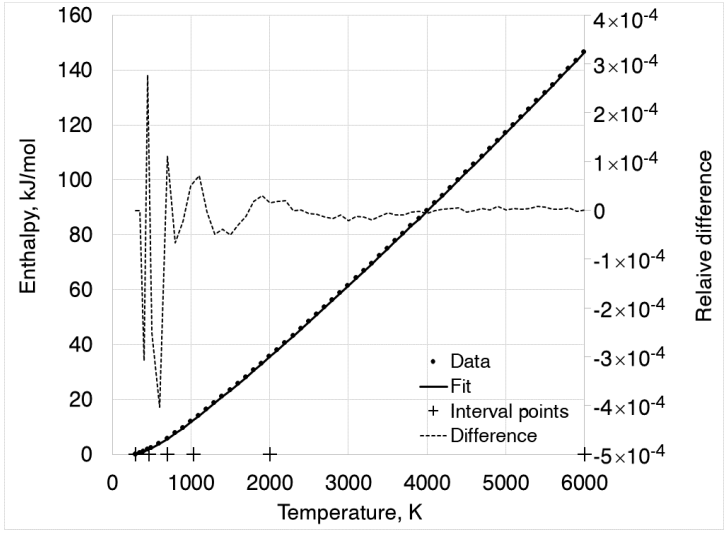


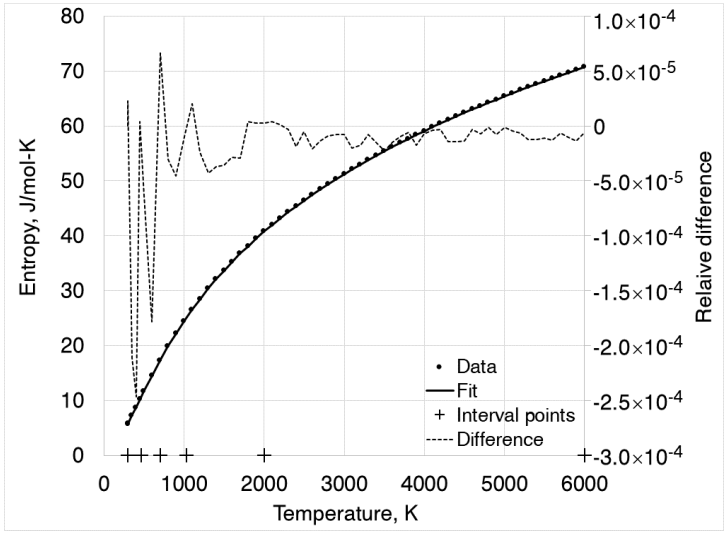


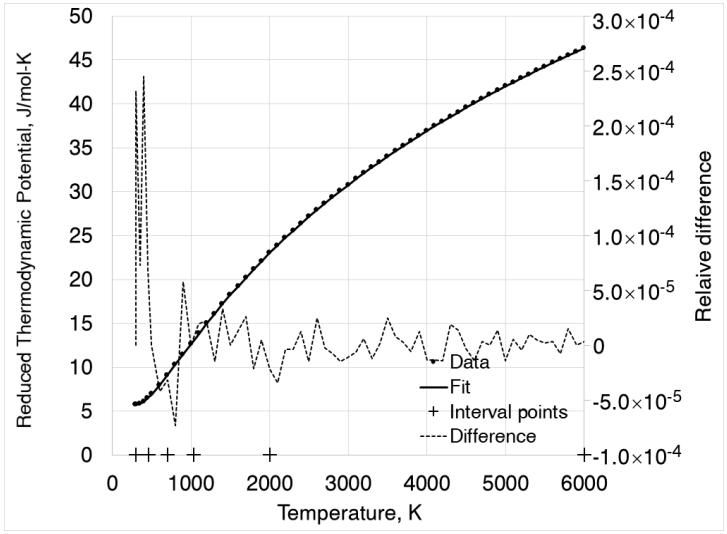
Carbon (C) – C1(ref)

MW = 12.011 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 5.74 J/mol-K – p^o = 0.1 MPa



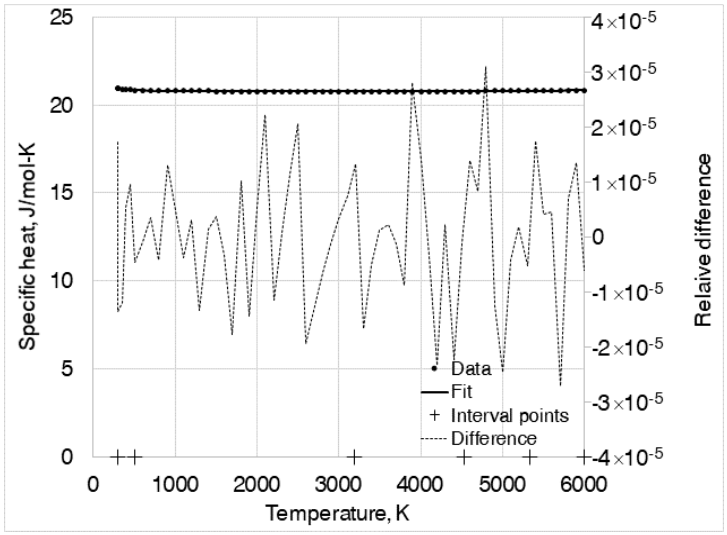


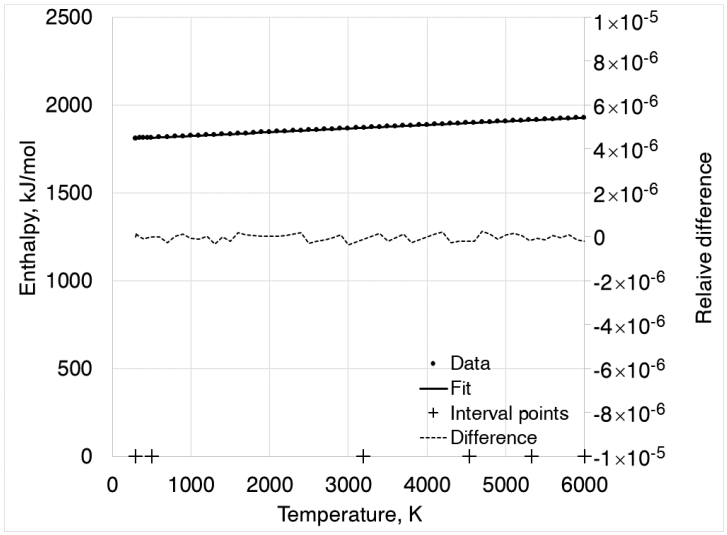


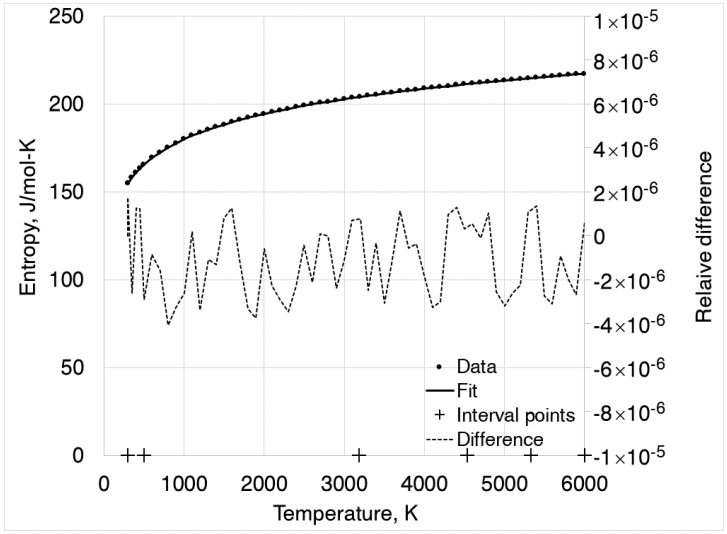


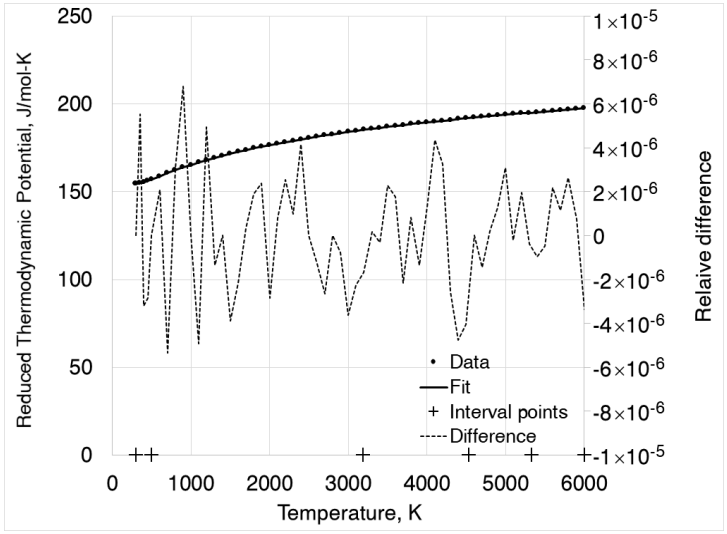
Carbon, Ion (C^+) – C1+(g)

MW = 12.01 g/mol – $\Delta_f H^o$(298.15 K) = 1809.435 kJ/mol – S^o(298.15 K) = 154.662 J/mol-K – p^o = 0.1 MPa



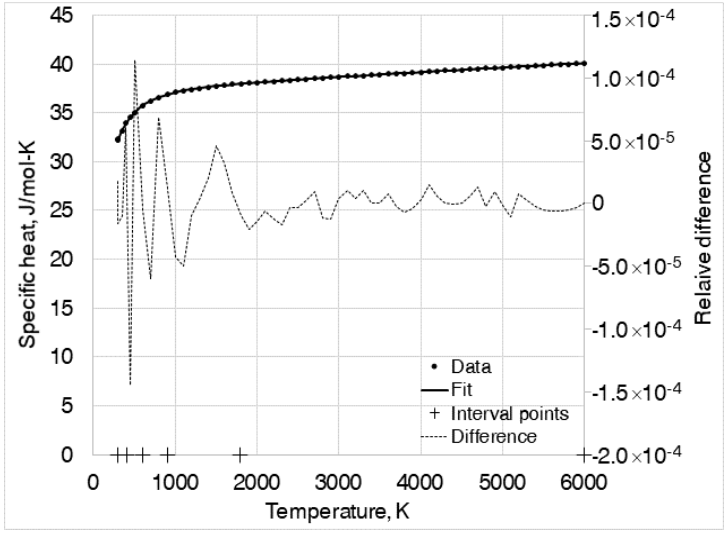


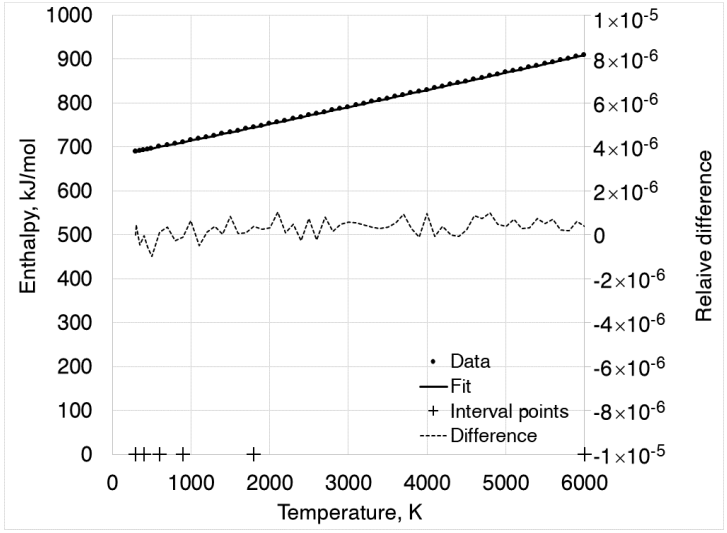


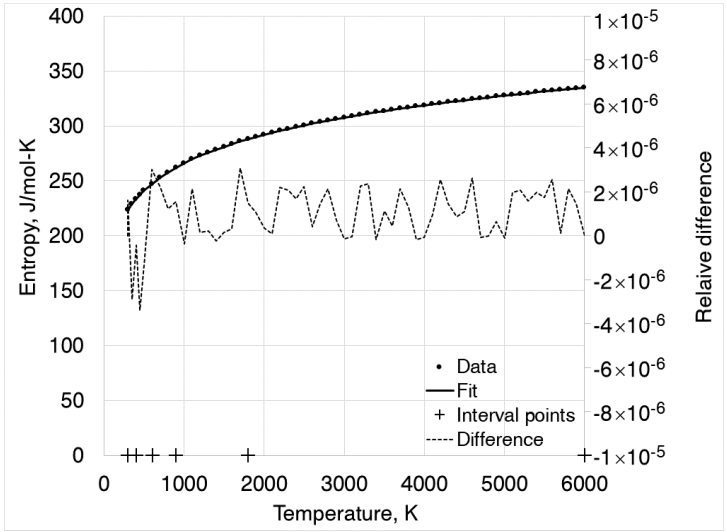


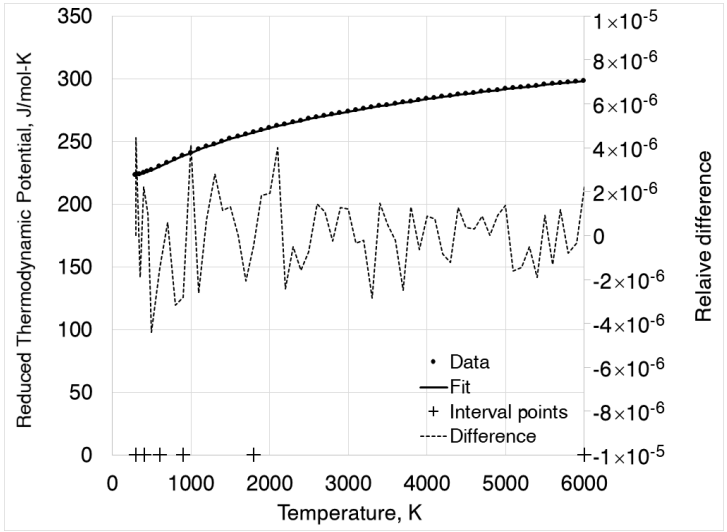
Aluminum Carbide (AlC) – C1Al1(g)

MW = 38.993 g/mol – $\Delta_f H^o$(298.15 K) = 689.523 kJ/mol – S^o(298.15 K) = 223.449 J/mol-K – p^o = 0.1 MPa



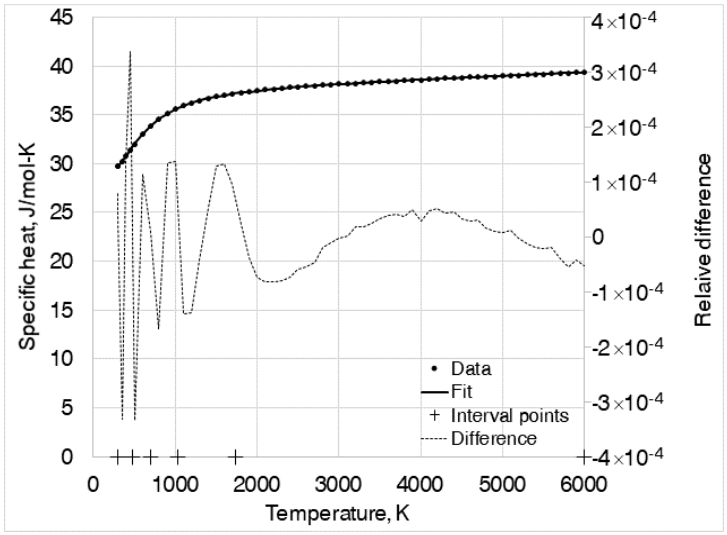


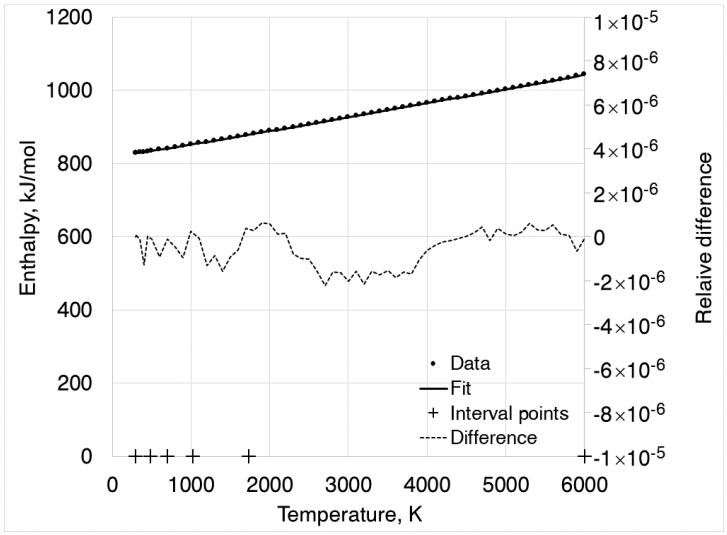


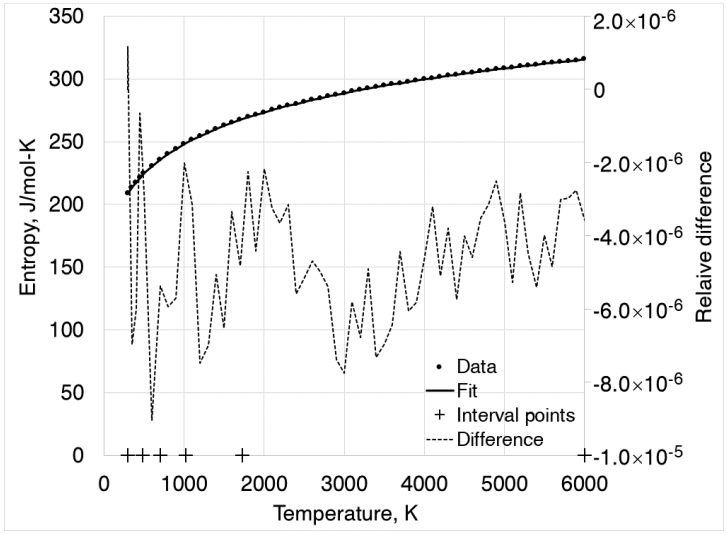


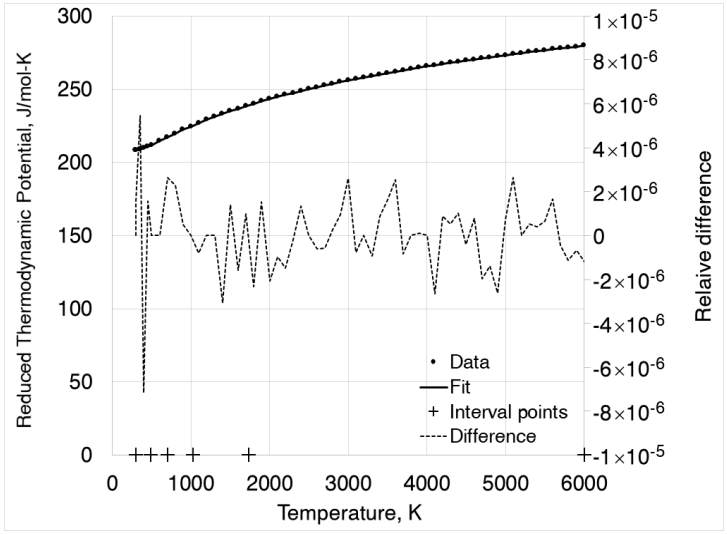
Boron Carbide (BC) – C1B1(g)

MW = 22.82 g/mol – $\Delta_f H^o$(298.15 K) = 828.432 kJ/mol – S^o(298.15 K) = 208.555 J/mol-K – p^o = 0.1 MPa



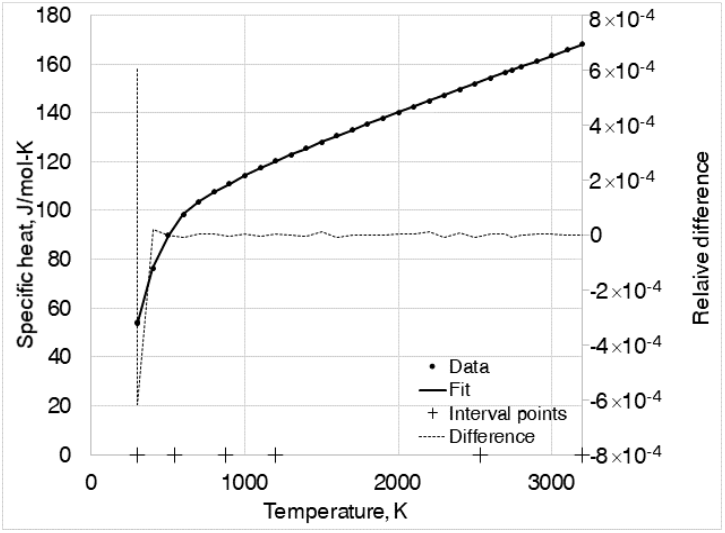


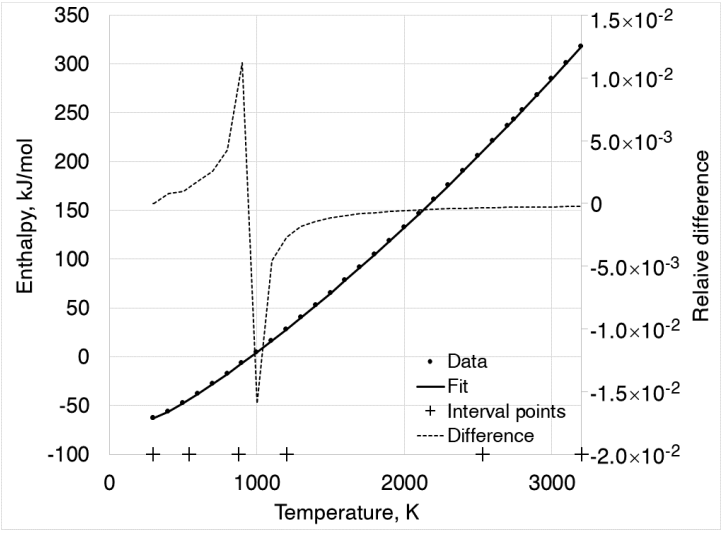


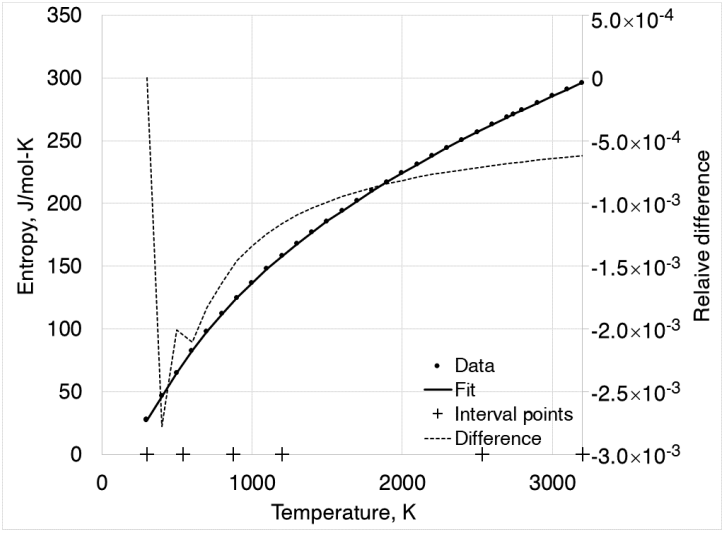


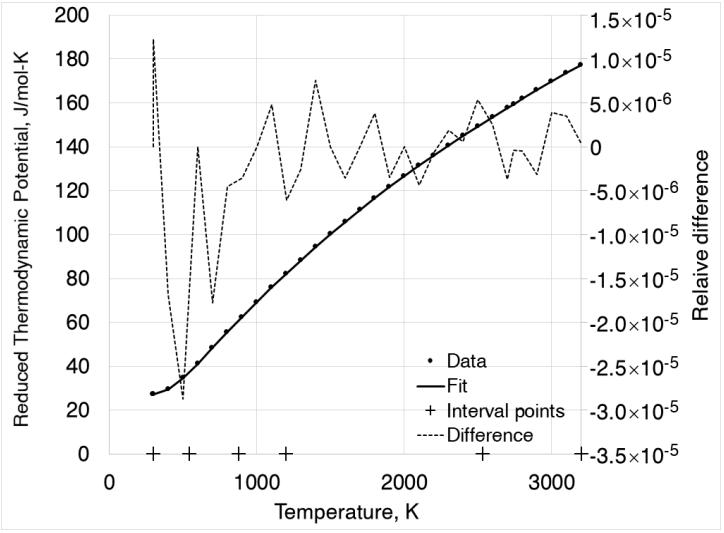
Boron Carbide (B_4C) – C1B4(cr)

MW = 55.25 g/mol – $\Delta_f H^o$(298.15 K) = -62.676 kJ/mol – S^o(298.15 K) = 27.175 J/mol-K – p^o = 0.1 MPa



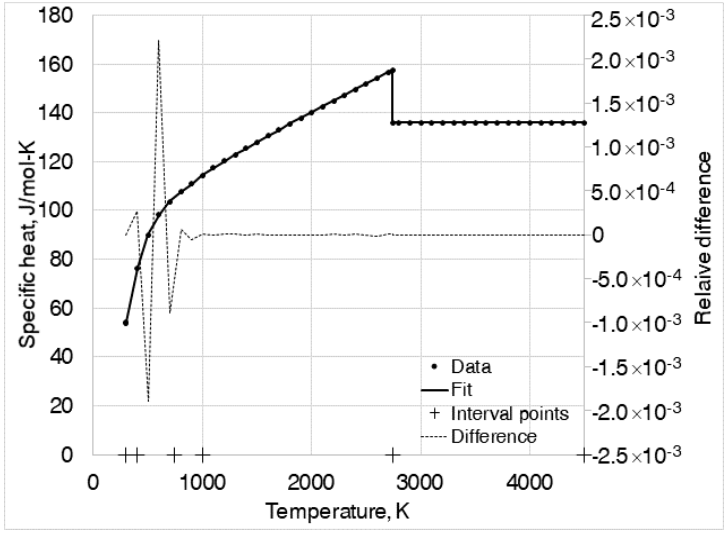


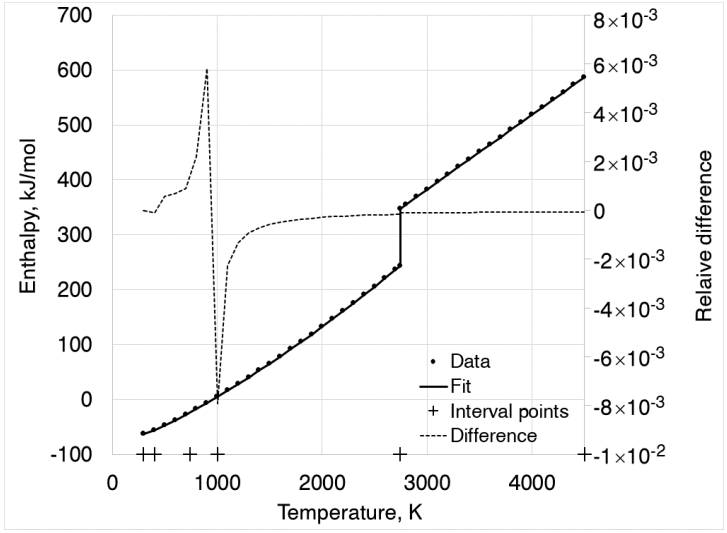


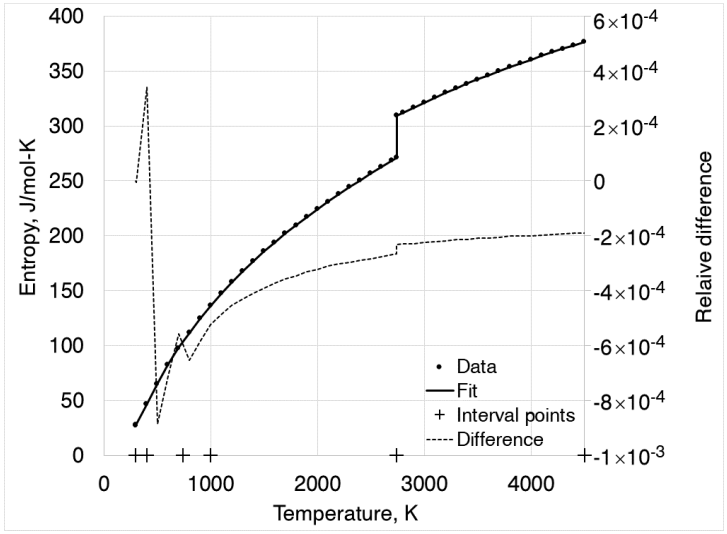


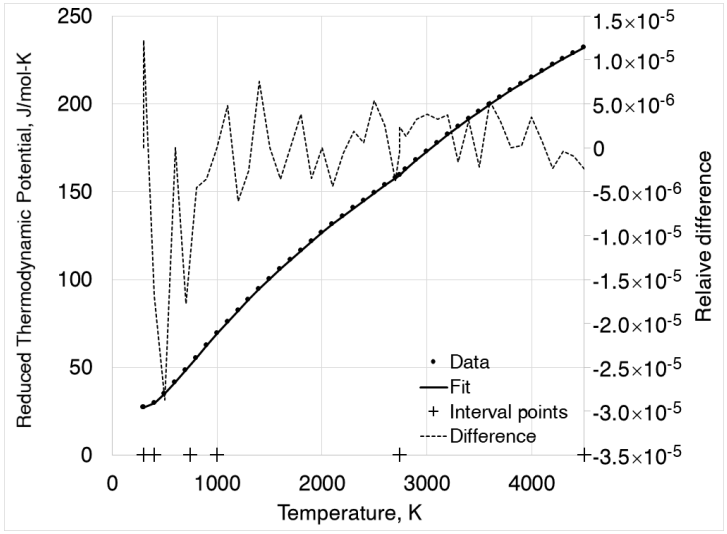
Boron Carbide (B_4C) – C1B4(cr,l)

MW = 55.25 g/mol – $\Delta_f H^o$(298.15 K) = -62.676 kJ/mol – S^o(298.15 K) = 27.175 J/mol-K – p^o = 0.1 MPa



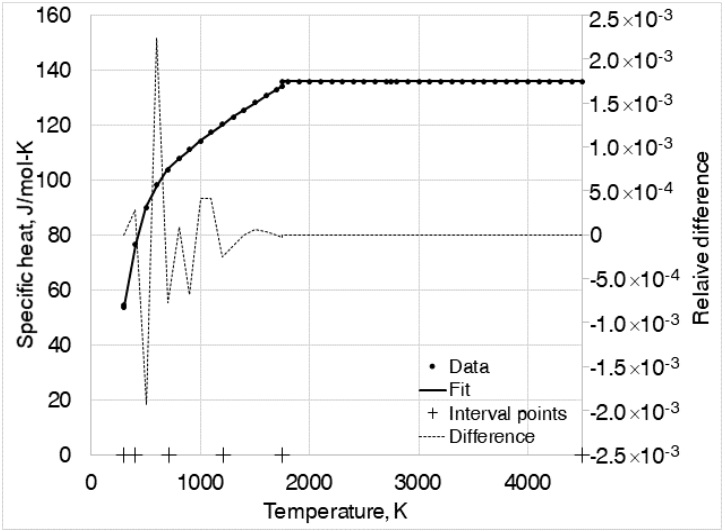


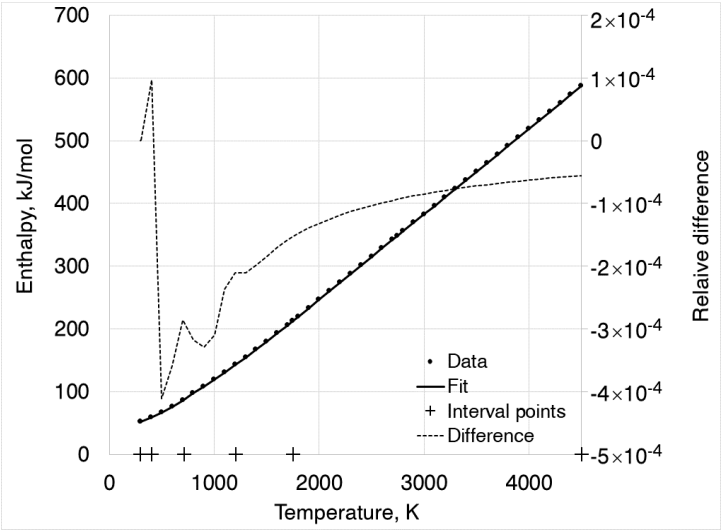


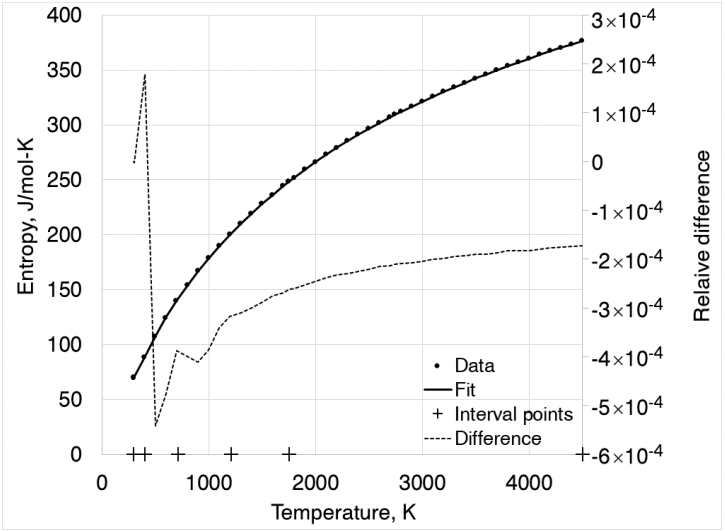


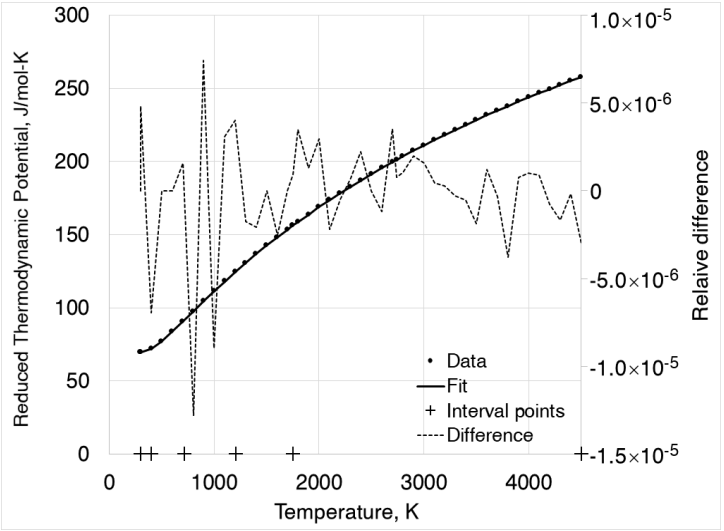
Boron Carbide (B_4C) – C1B4(l)

MW = 55.25 g/mol – $\Delta_f H^o$(298.15 K) = 51.885 kJ/mol – S^o(298.15 K) = 69.426 J/mol-K – p^o = 0.1 MPa



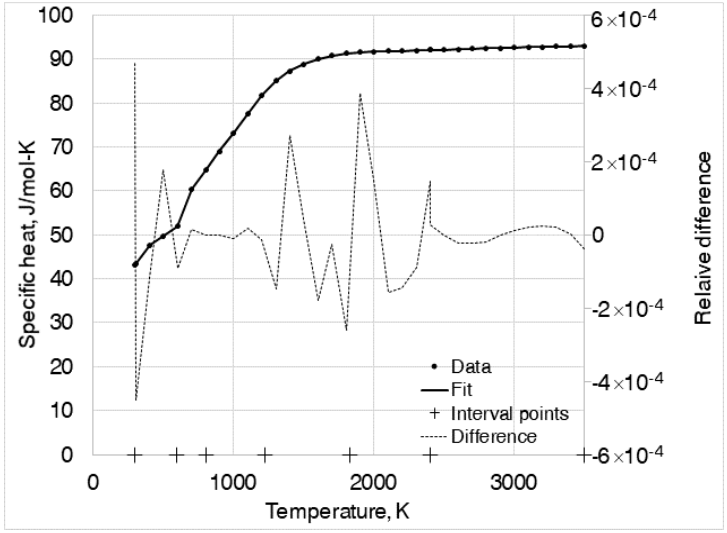


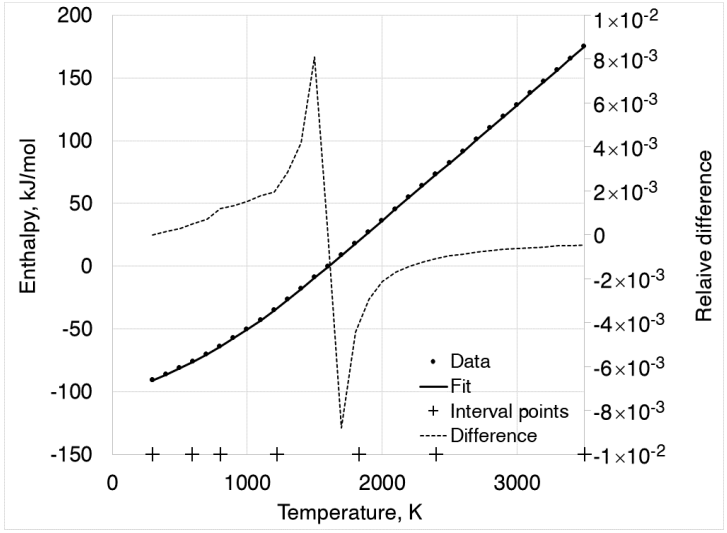


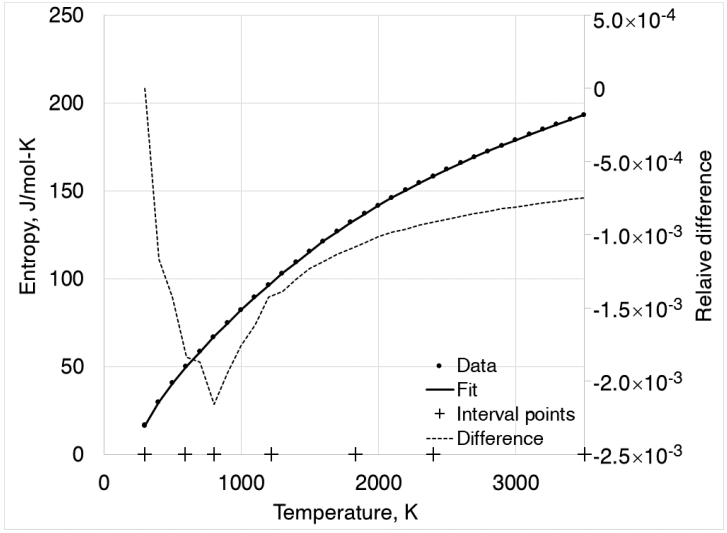


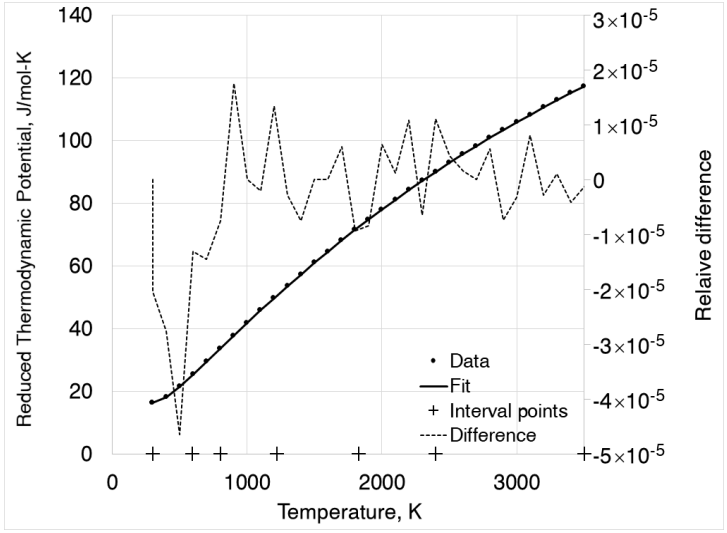
Beryllium Carbide (Be_2C) – C1Be2(cr)

MW = 30.035 g/mol – $\Delta_f H^o$(298.15 K) = -90.793 kJ/mol – S^o(298.15 K) = 16.318 J/mol-K – p^o = 0.1 MPa



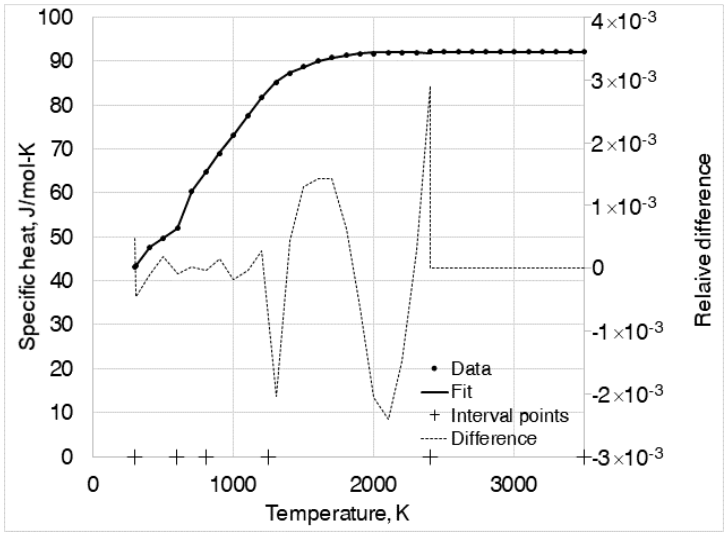


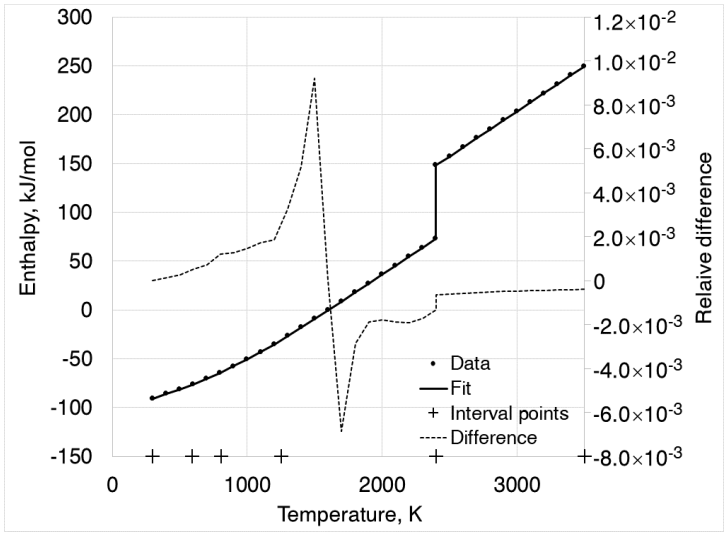


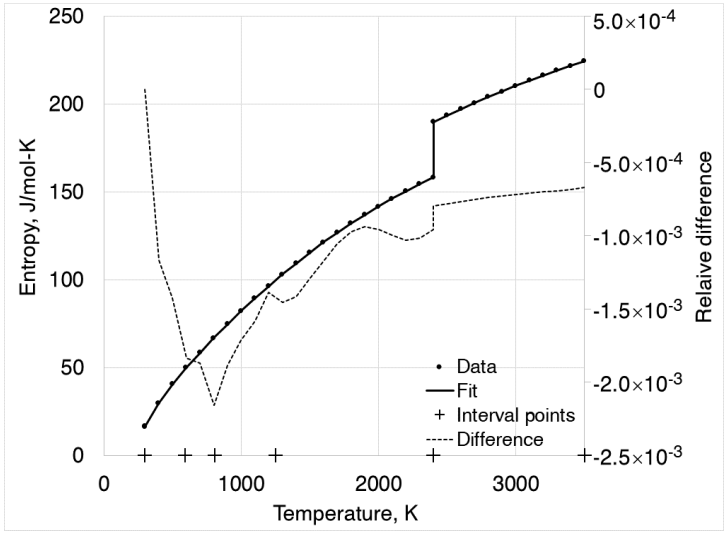


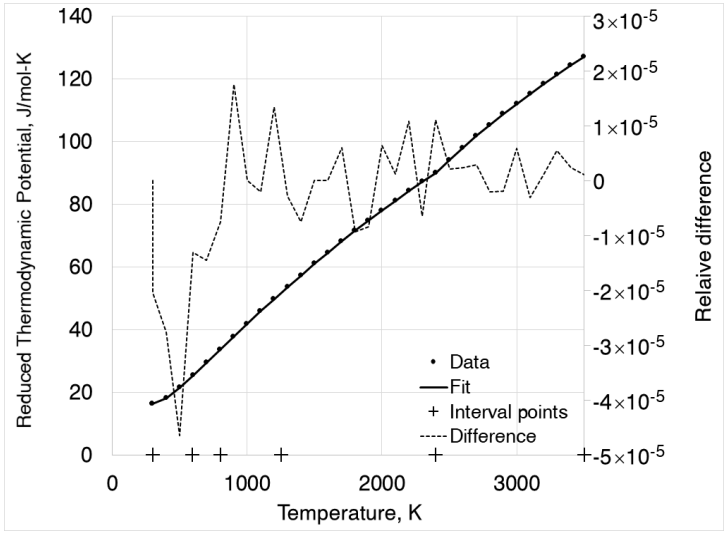
Beryllium Carbide (Be_2C) – C1Be2(cr,l)

MW = 30.035 g/mol – $\Delta_f H^o$(298.15 K) = -90.793 kJ/mol – S^o(298.15 K) = 16.318 J/mol-K – p^o = 0.1 MPa



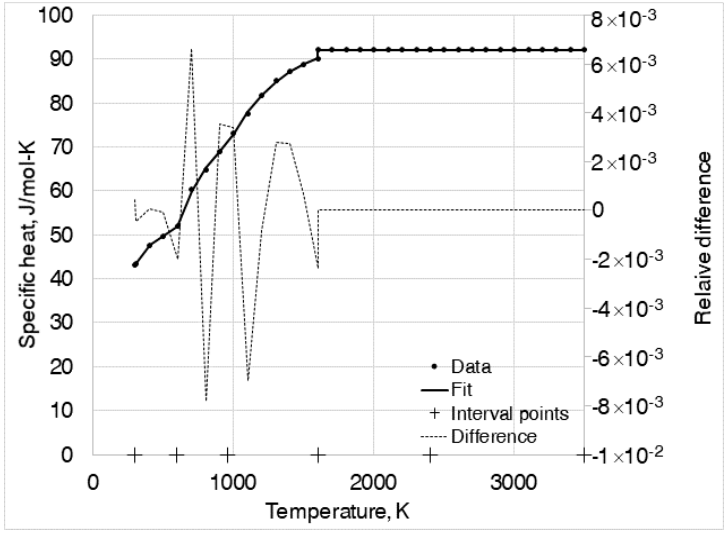


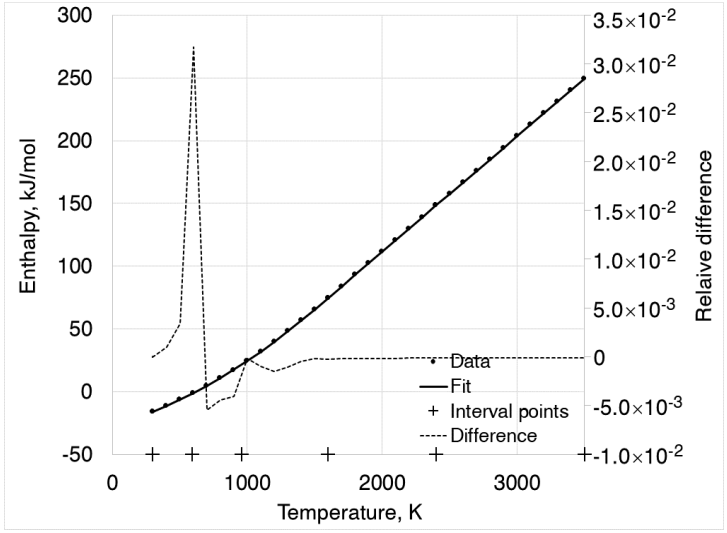


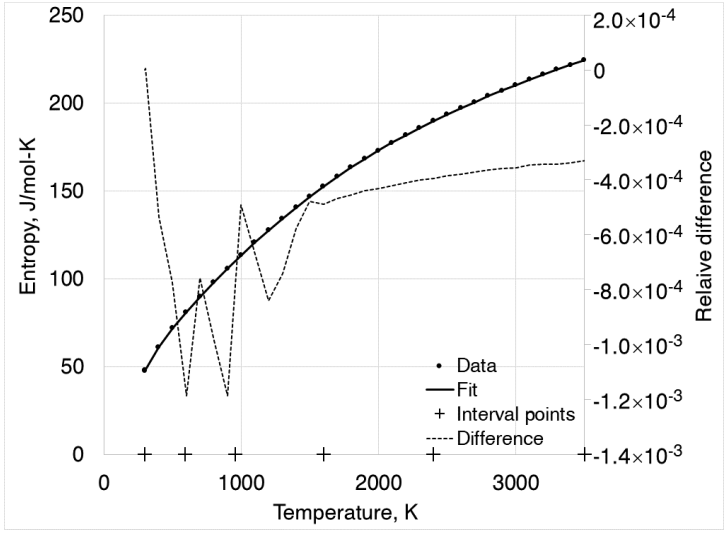


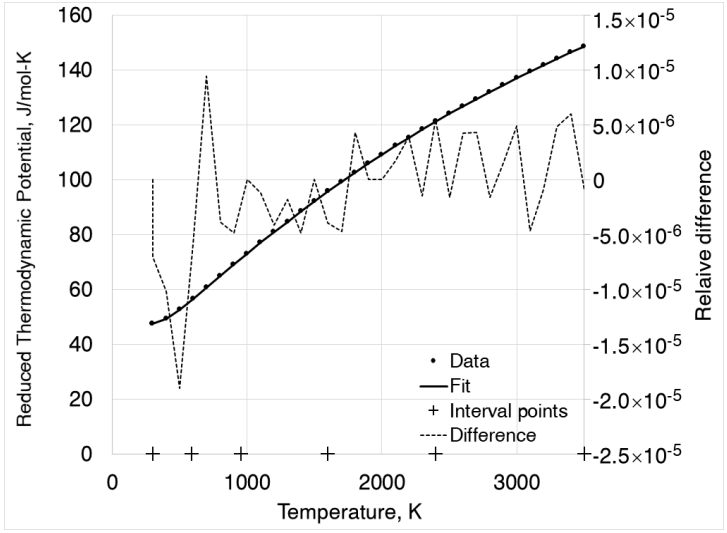
Beryllium Carbide (Be_2C) – C1Be2(l)

MW = 30.035 g/mol – $\Delta_f H^o$(298.15 K) = -15.893 kJ/mol – S^o(298.15 K) = 47.467 J/mol-K – p^o = 0.1 MPa



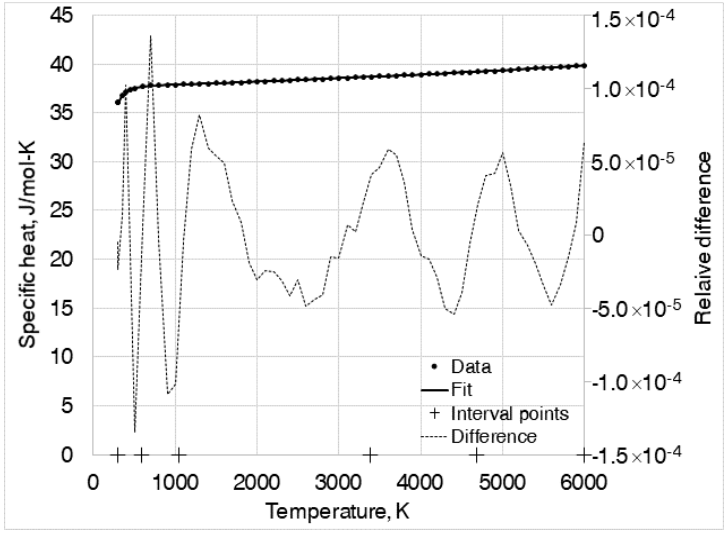


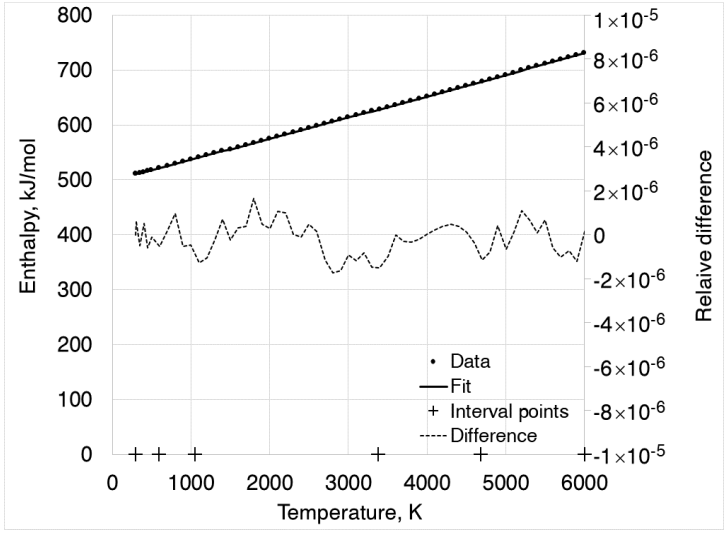


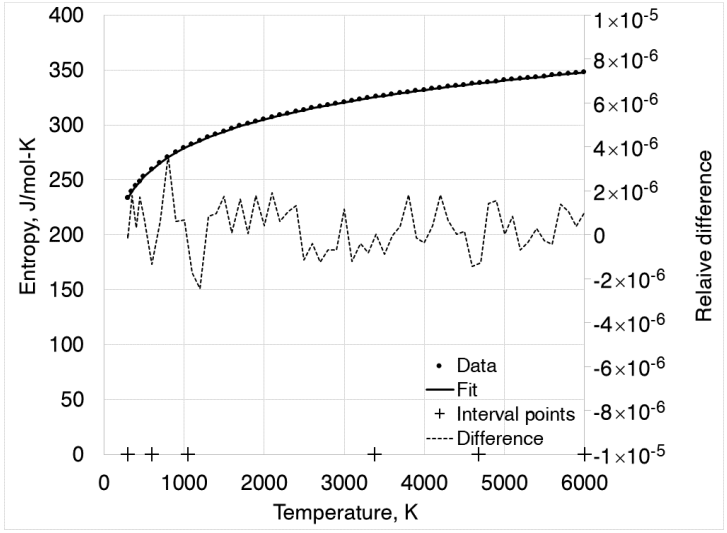


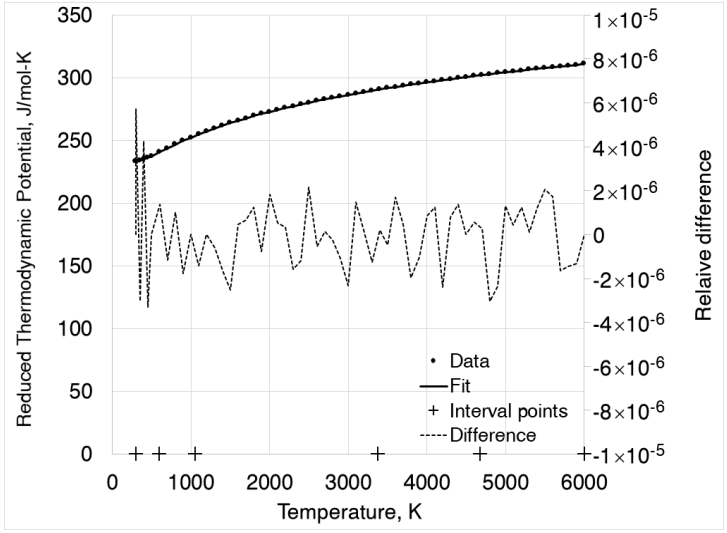
Bromomethylidyne (CBr) – C1Br1(g)

MW = 91.915 g/mol – $\Delta_f H^o$(298.15 K) = 510.448 kJ/mol – S^o(298.15 K) = 233.452 J/mol-K – p^o = 0.1 MPa



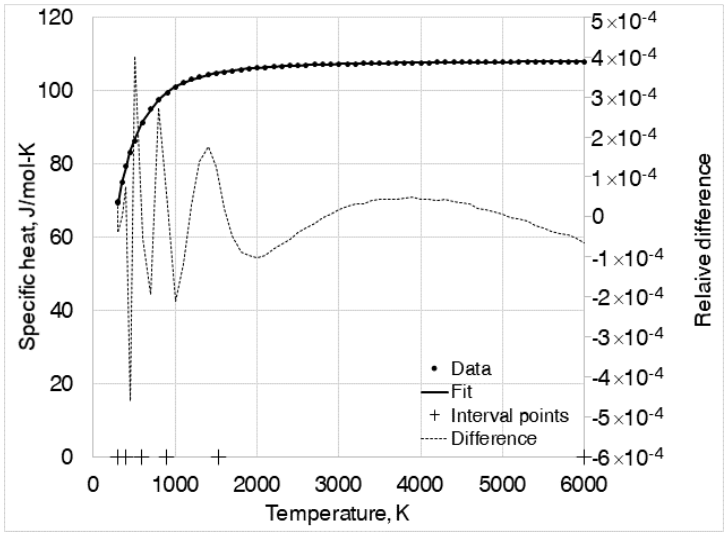


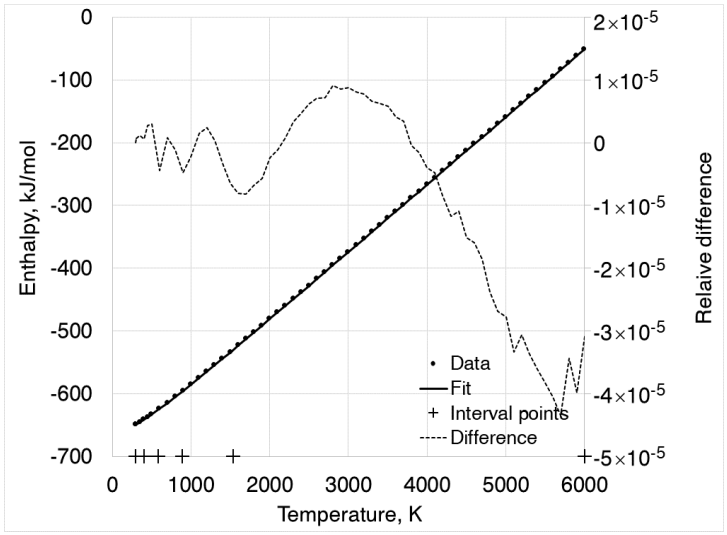


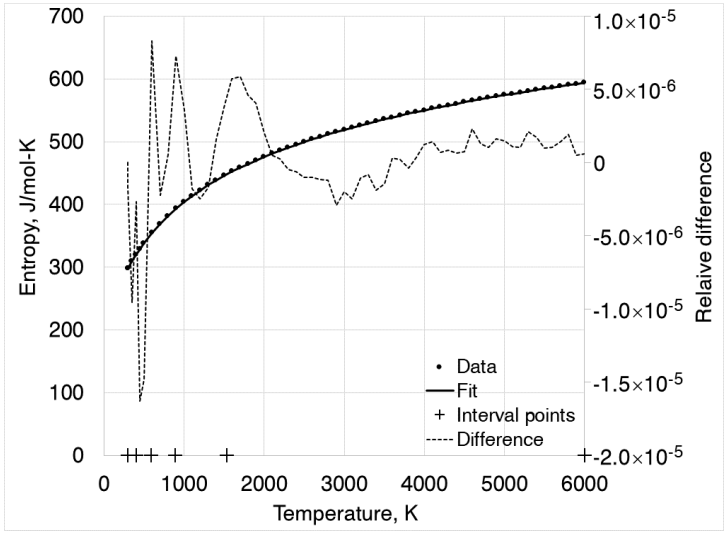


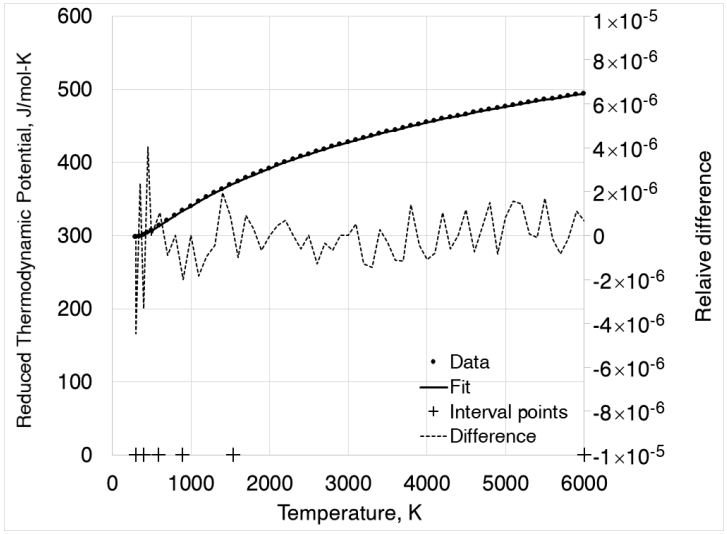
Bromotrifluoromethane ($CBrF_3$) – C1Br1F3(g)

MW = 148.91 g/mol – $\Delta_f H^o$(298.15 K) = -648.938 kJ/mol – S^o(298.15 K) = 297.827 J/mol-K – p^o = 0.1 MPa



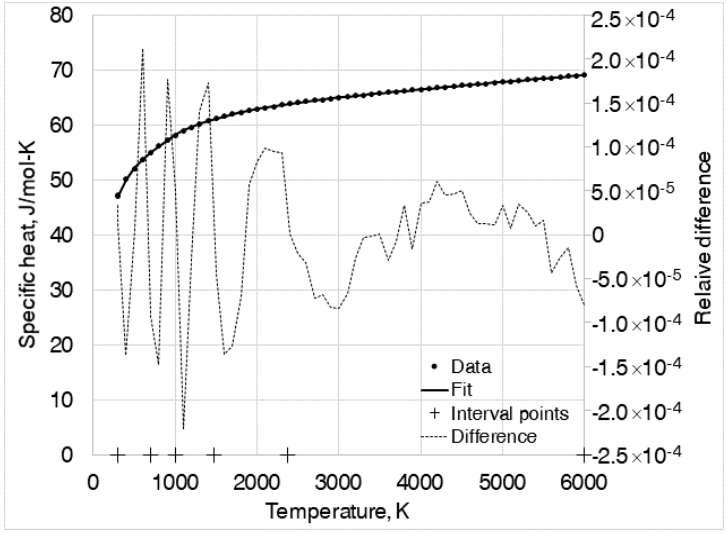


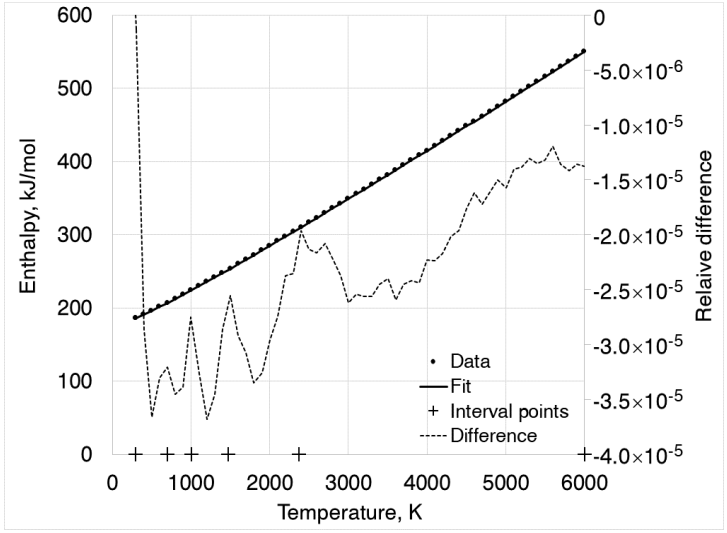


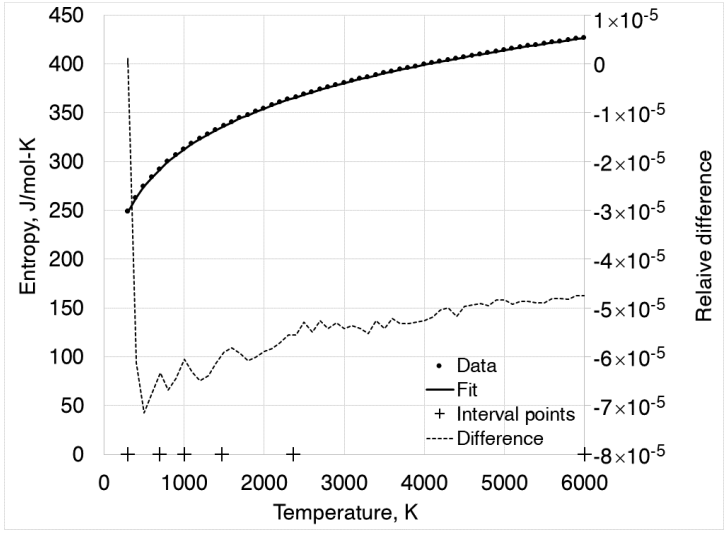


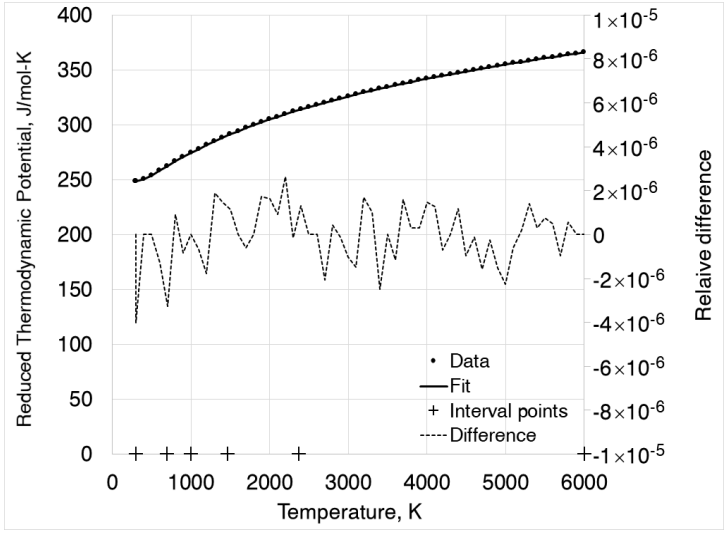
Cyanogen Bromide (BrCN) – C1Br1N1(g)

MW = 105.922 g/mol – $\Delta_f H^o$(298.15 K) = 186.188 kJ/mol – S^o(298.15 K) = 248.363 J/mol-K – p^o = 0.1 MPa



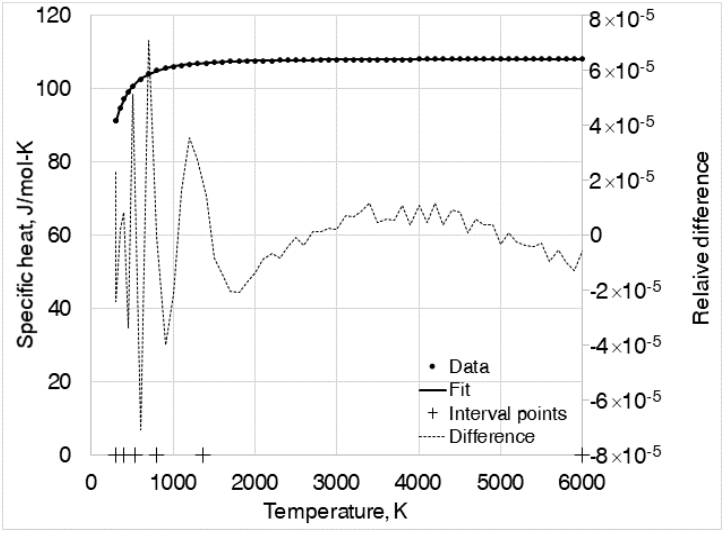


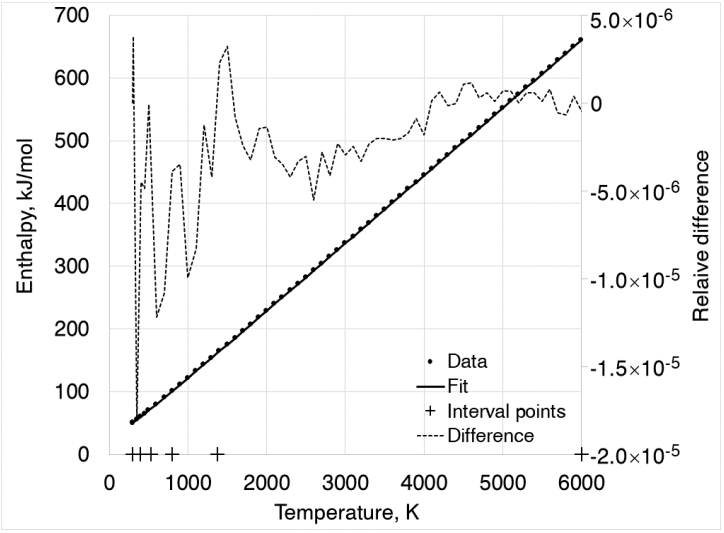


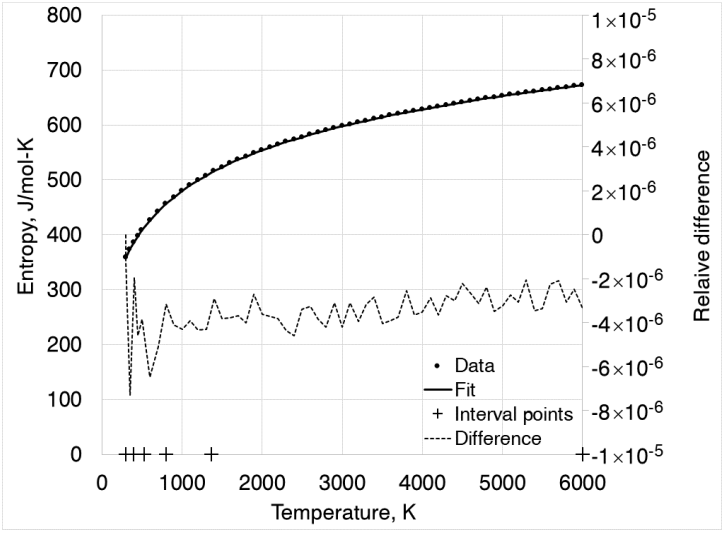


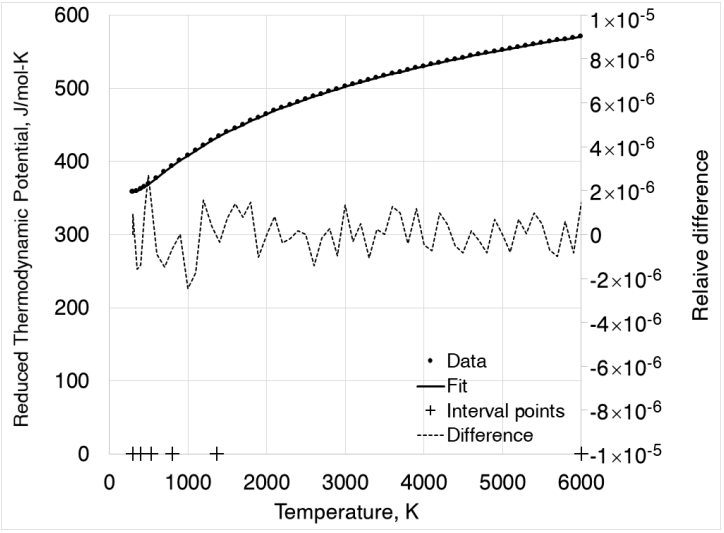
Tetrabromomethane (CBr_4) – C1Br4(g)

MW = 331.627 g/mol – $\Delta_f H^o$(298.15 K) = 50.208 kJ/mol – S^o(298.15 K) = 358.095 J/mol-K – p^o = 0.1 MPa



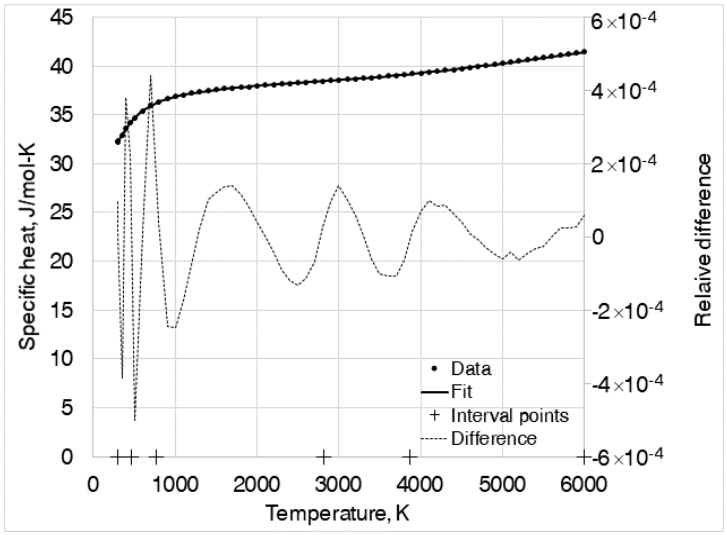


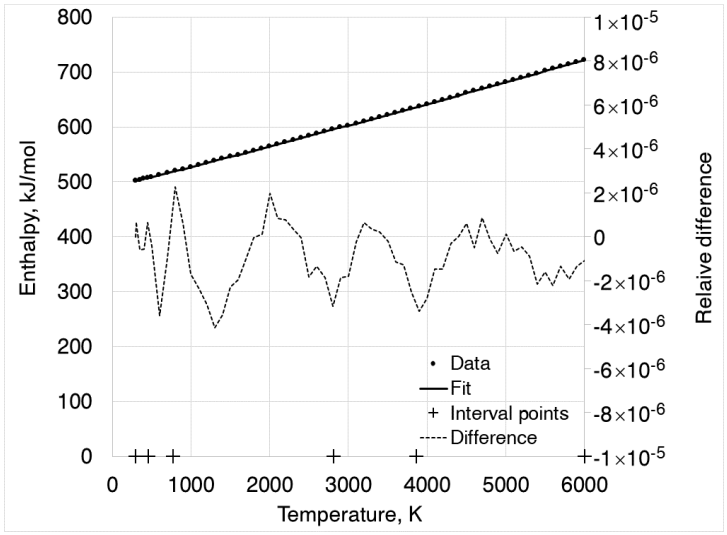


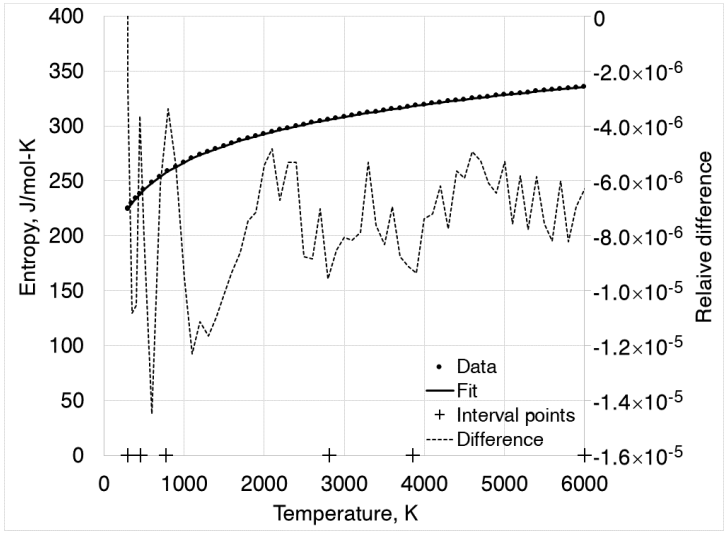


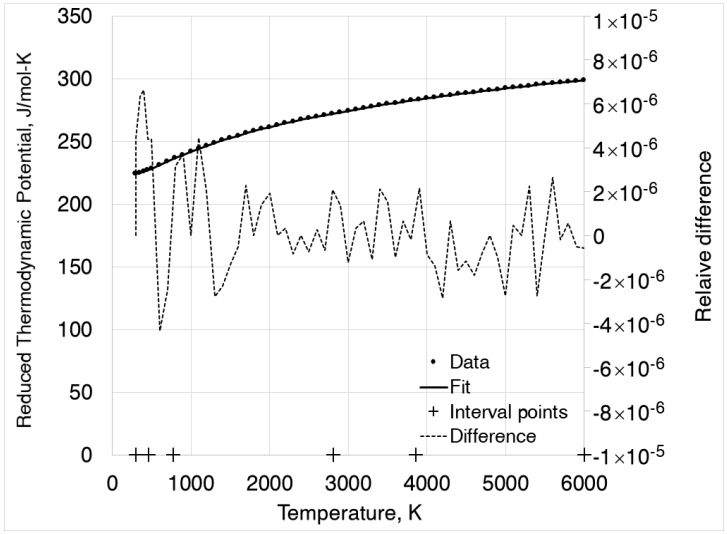
Chloromethylidyne (CCl) – C1Cl1(g)

MW = 47.464 g/mol – $\Delta_f H^o$(298.15 K) = 502.08 kJ/mol – S^o(298.15 K) = 224.516 J/mol-K – p^o = 0.1 MPa



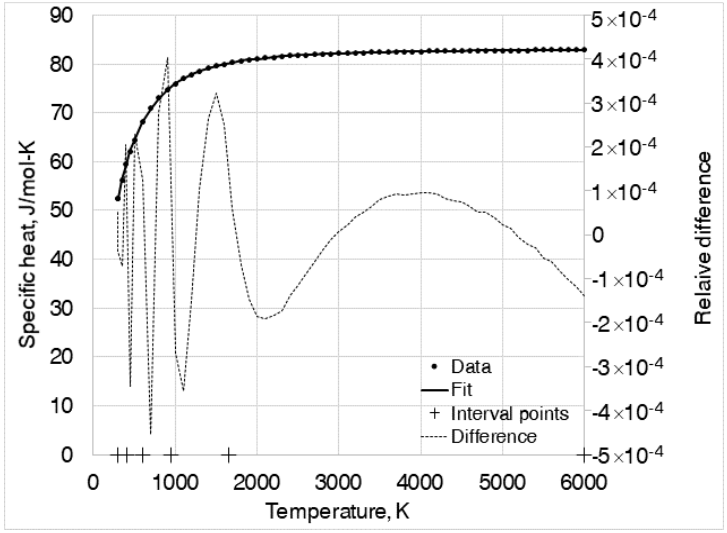


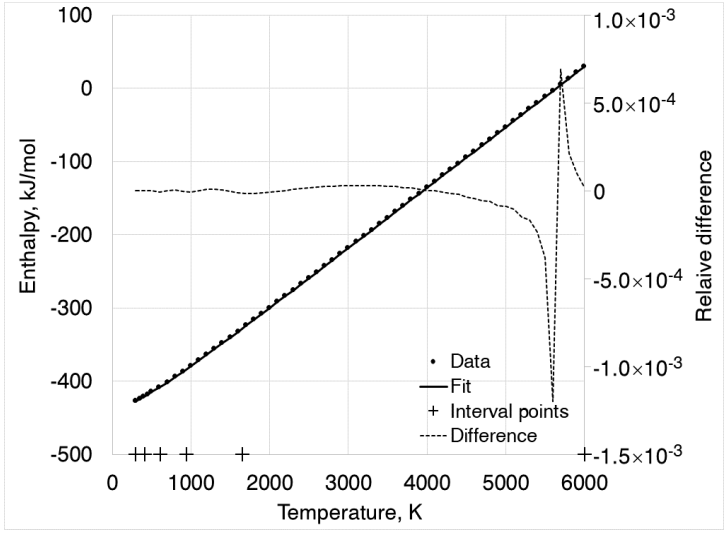


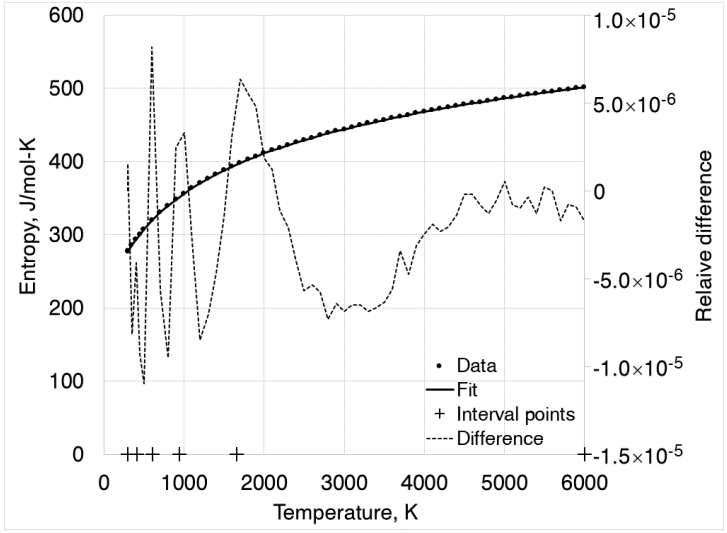


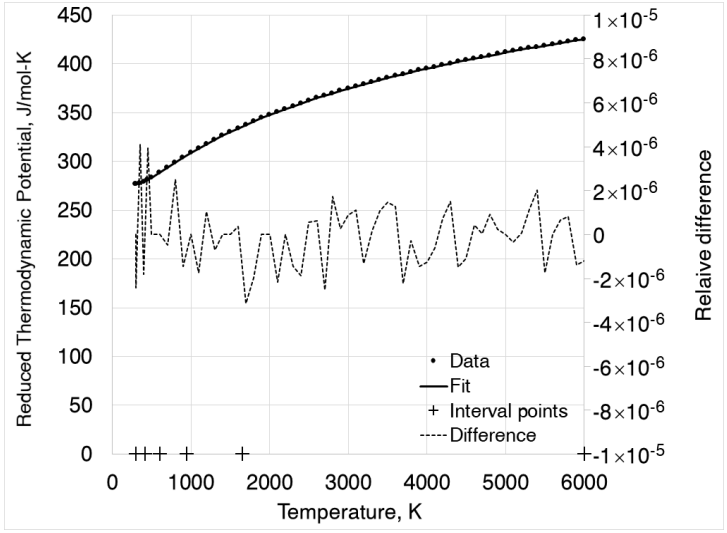
Carbonic Chloride Fluoride (COClF) – C1Cl1F1O1(g)

MW = 82.462 g/mol – $\Delta_f H^o$(298.15 K) = -426.768 kJ/mol – S^o(298.15 K) = 277.012 J/mol-K – p^o = 0.1 MPa



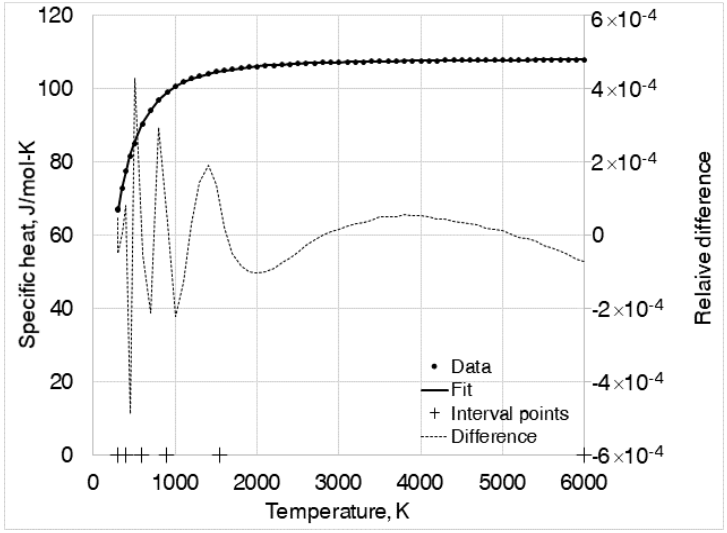


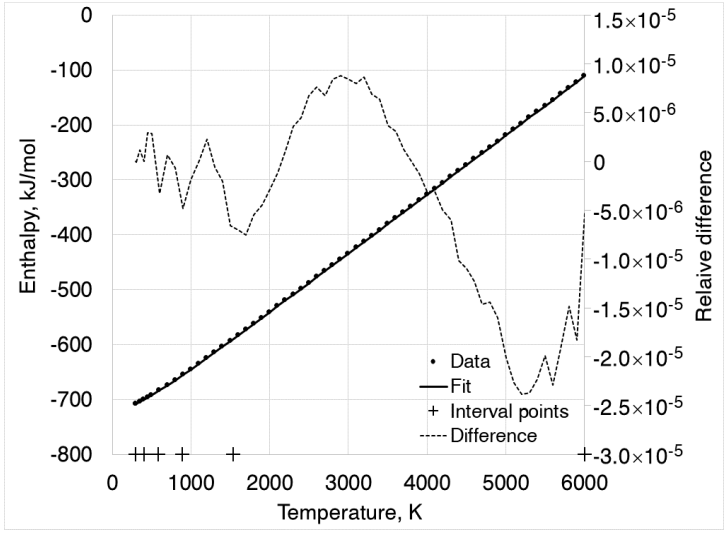


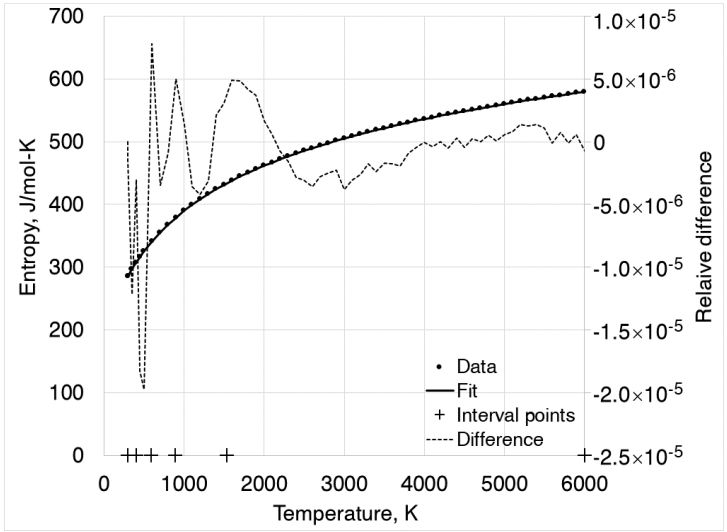


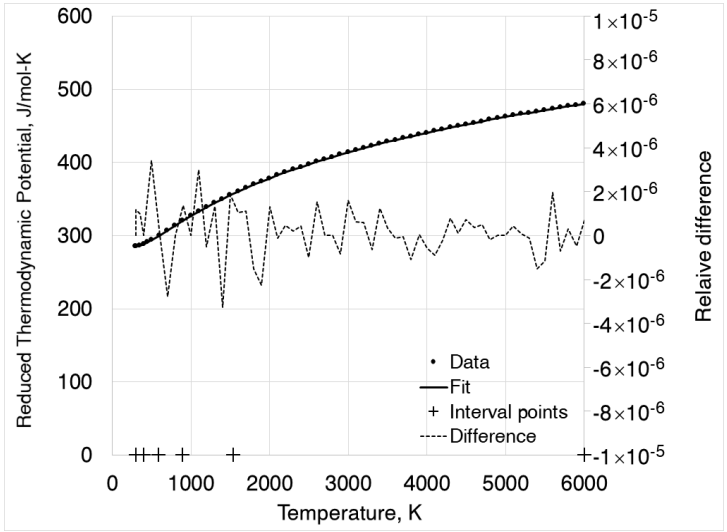
Chlorotrifluoromethane ($CClF_3$) – C1Cl1F3(g)

MW = 104.459 g/mol – $\Delta_f H^o$(298.15 K) = -707.933 kJ/mol – S^o(298.15 K) = 285.353 J/mol-K – p^o = 0.1 MPa



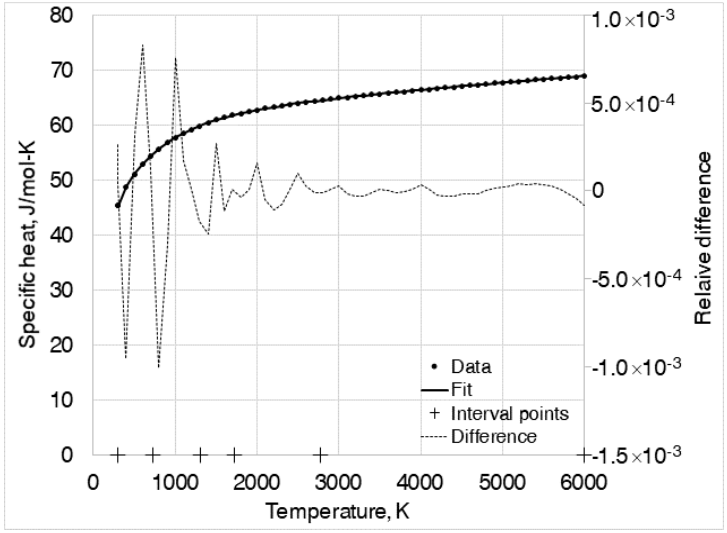


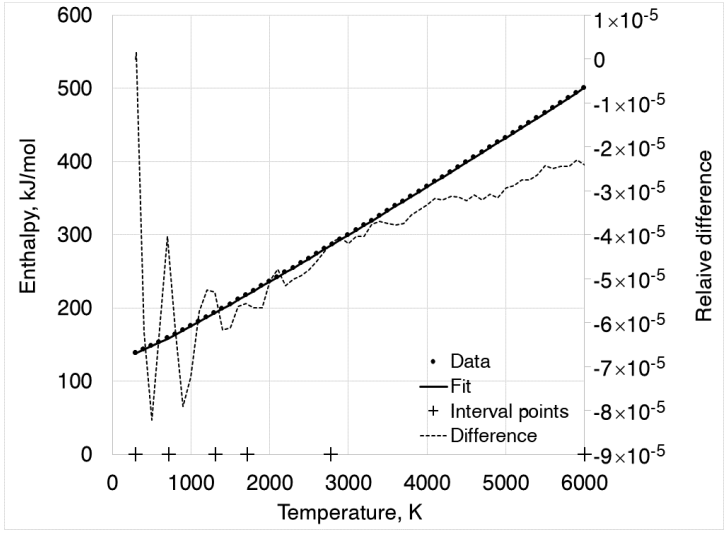


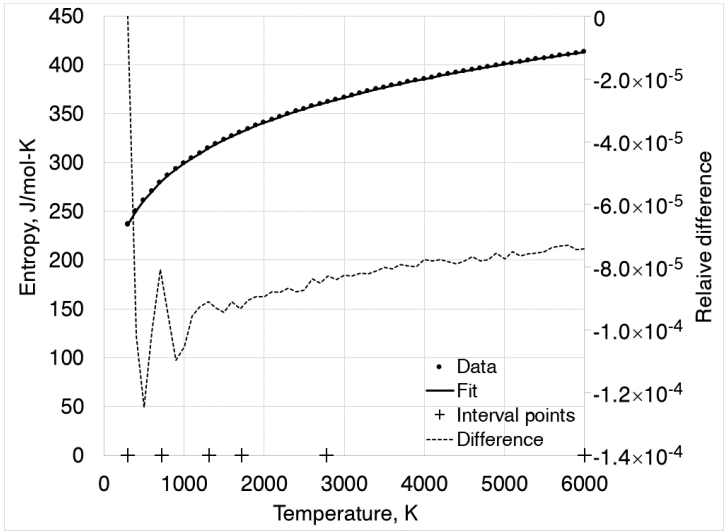


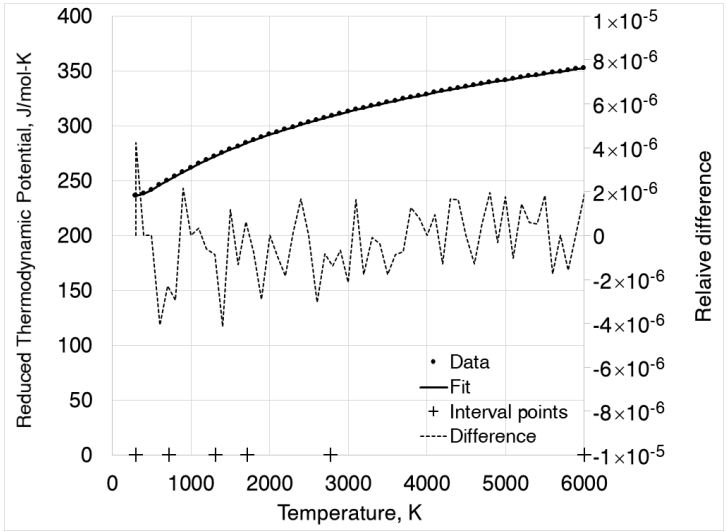
Cyanogen Chloride (ClCN) – C1Cl1N1(g)

MW = 61.471 g/mol – $\Delta_f H^o$(298.15 K) = 137.946 kJ/mol – S^o(298.15 K) = 236.334 J/mol-K – p^o = 0.1 MPa



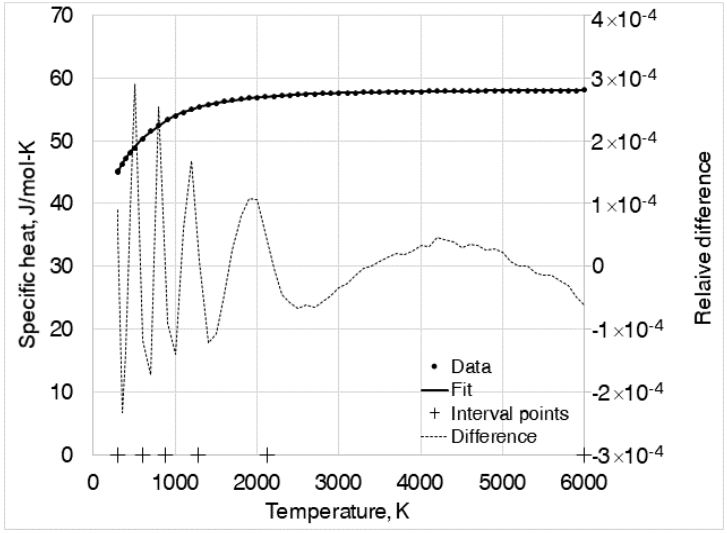


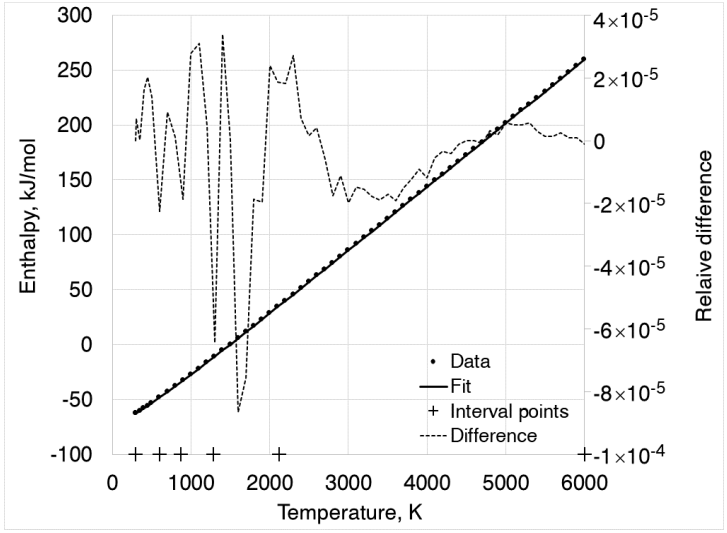


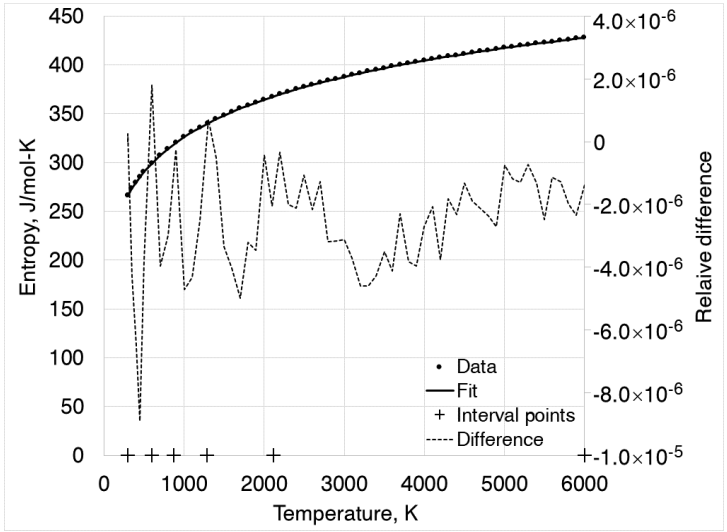


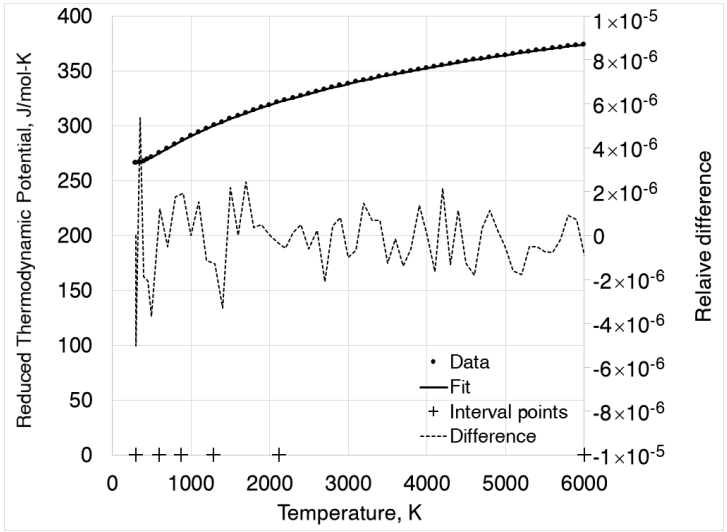
Carbonyl Chloride (COCl) – C1Cl1O1(g)

MW = 63.463 g/mol – $\Delta_f H^o$(298.15 K) = -62.76 kJ/mol – S^o(298.15 K) = 265.965 J/mol-K – p^o = 0.1 MPa



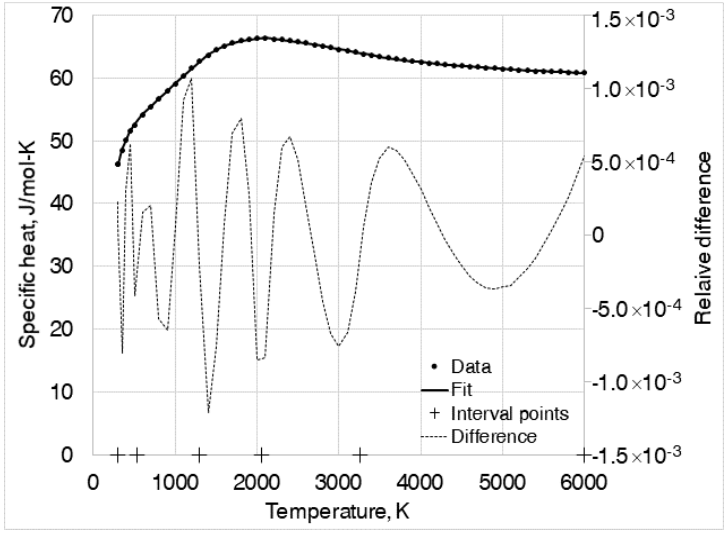


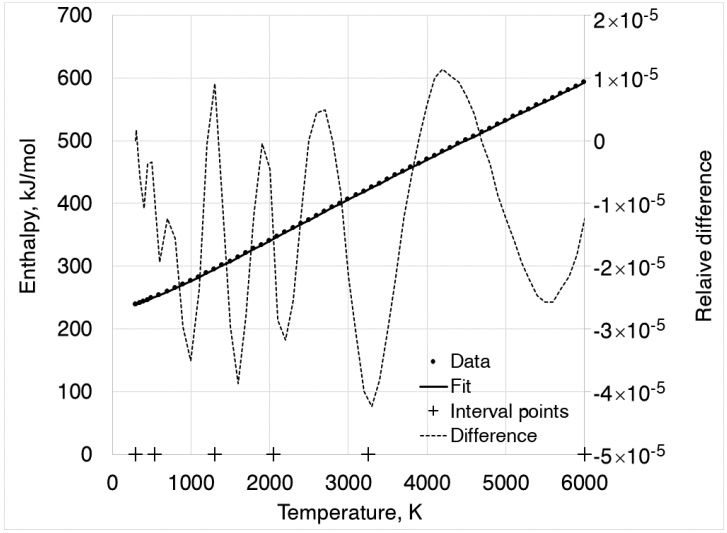


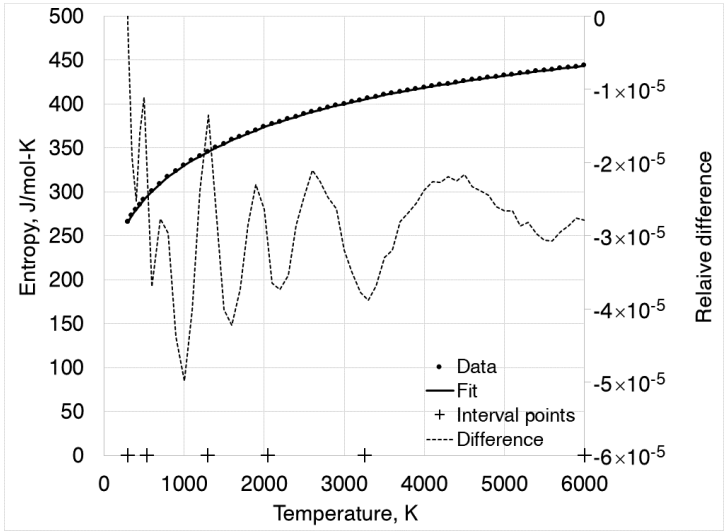


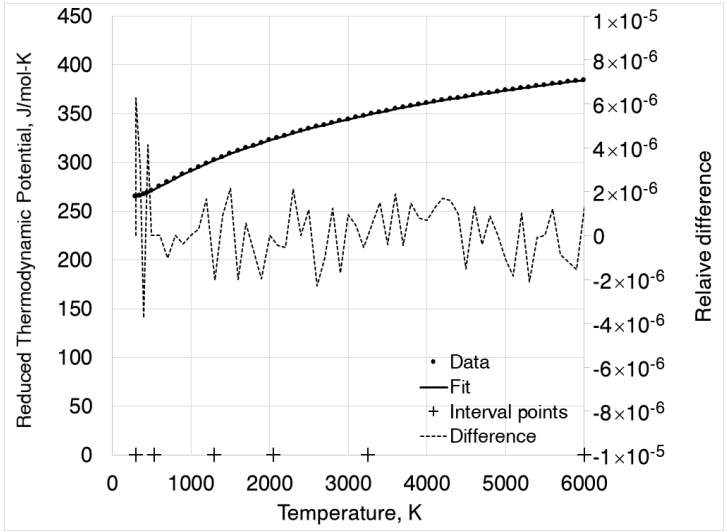
Dichloromethylene (CCl_2) – C1Cl2(g)

MW = 82.917 g/mol – $\Delta_f H^o$(298.15 K) = 238.488 kJ/mol – S^o(298.15 K) = 265.346 J/mol-K – p^o = 0.1 MPa



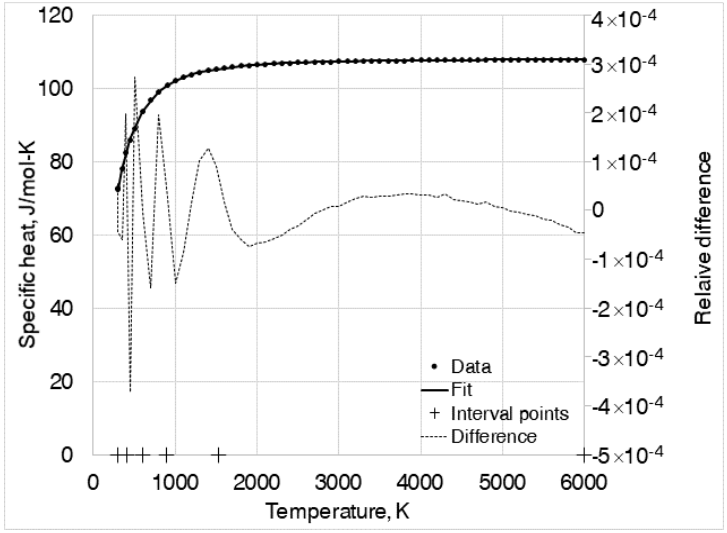


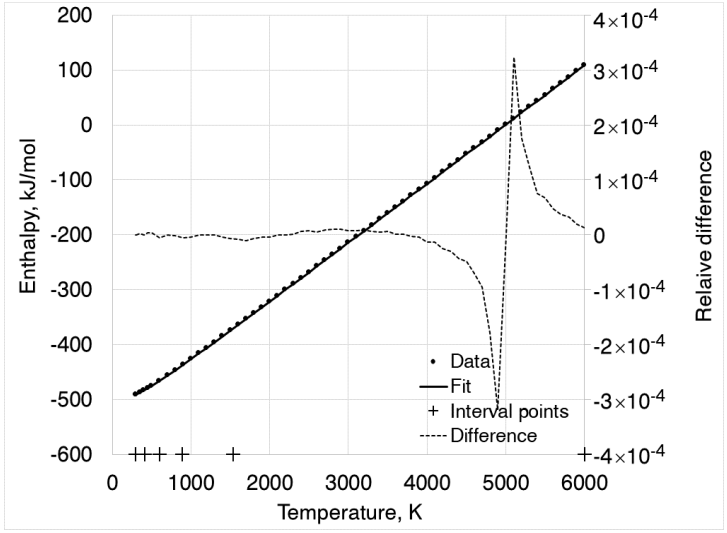


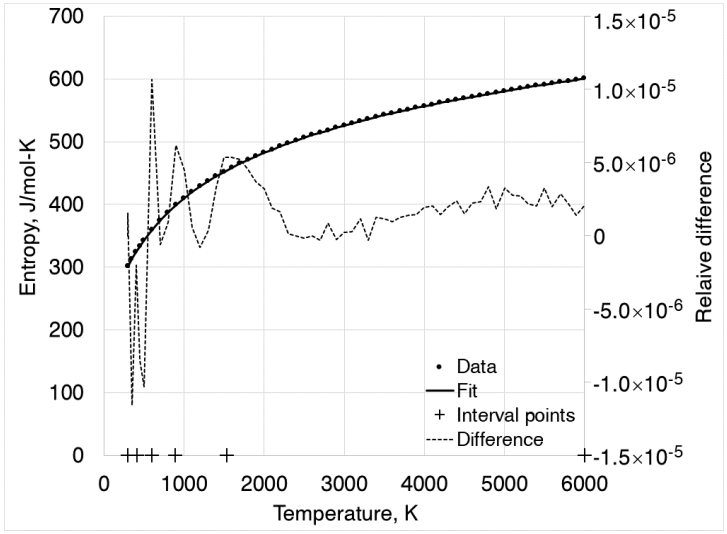


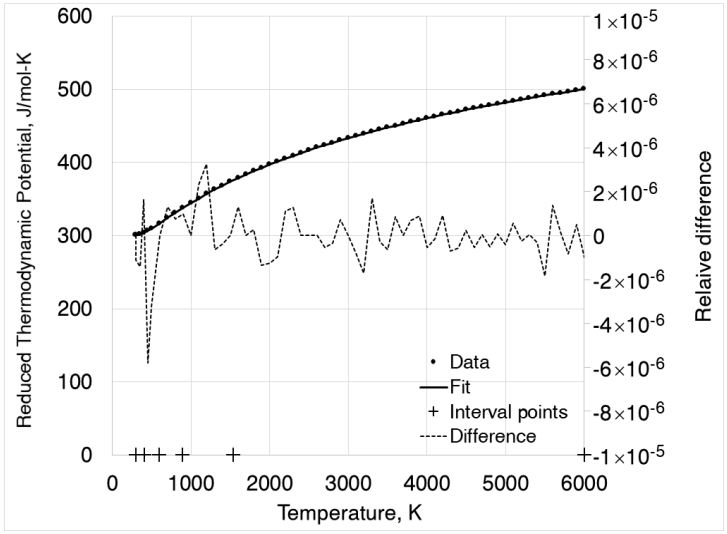
Dichlorodifluoromethane (CCl_2F_2) – C1Cl2F2(g)

MW = 120.914 g/mol – $\Delta_f H^o$(298.15 K) = -491.62 kJ/mol – S^o(298.15 K) = 300.897 J/mol-K – p^o = 0.1 MPa



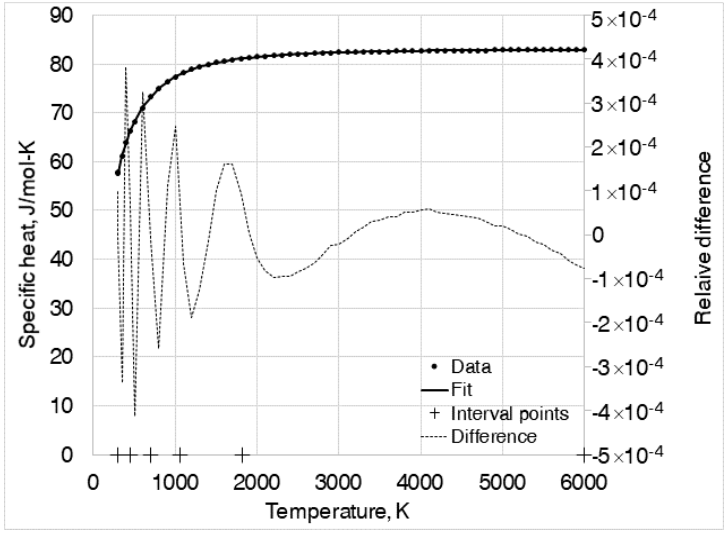


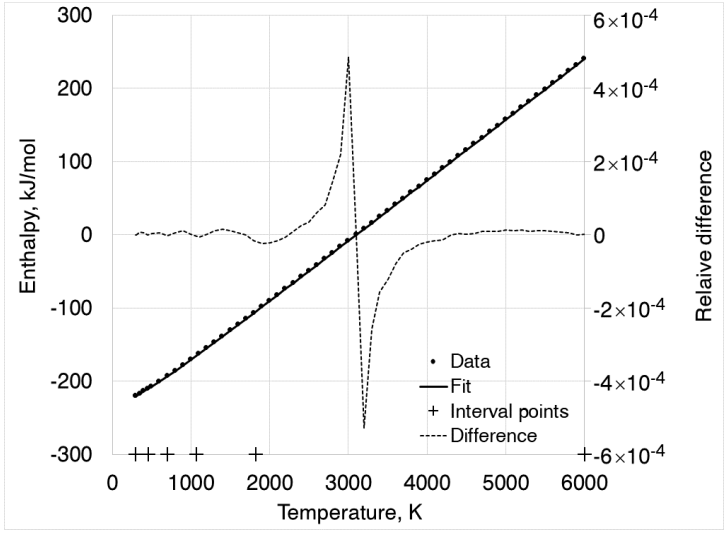


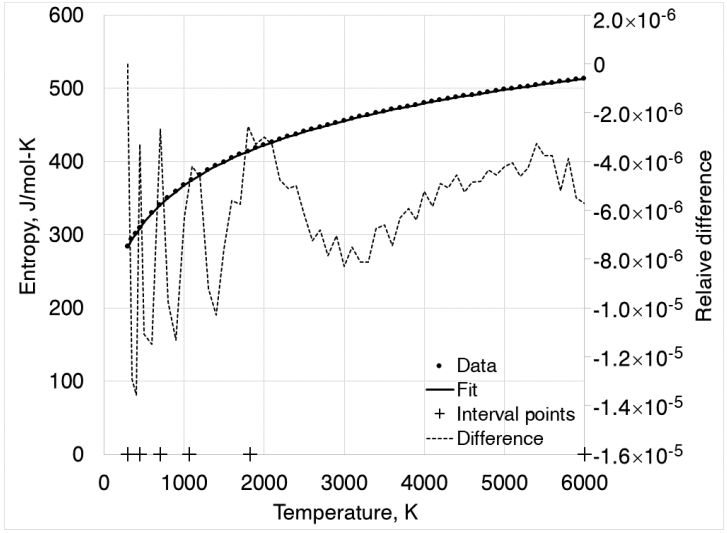


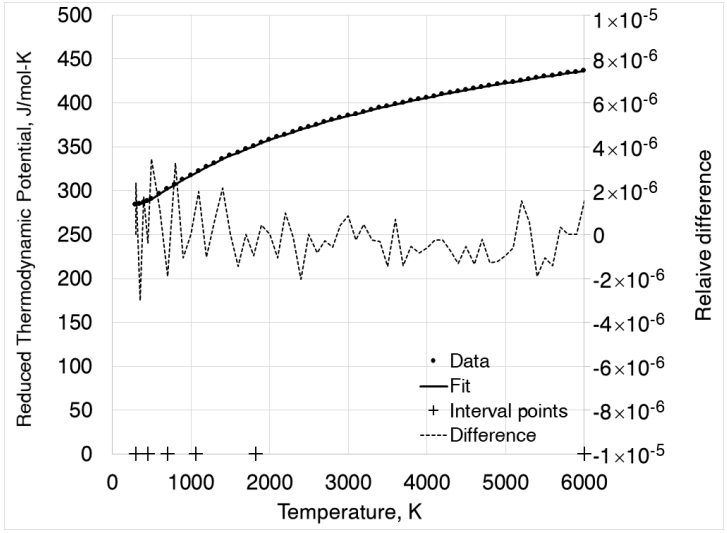
Carbonic Dichloride ($COCl_2$) – C1Cl2O1(g)

MW = 98.916 g/mol – $\Delta_f H^o$(298.15 K) = -220.078 kJ/mol – S^o(298.15 K) = 283.796 J/mol-K – p^o = 0.1 MPa



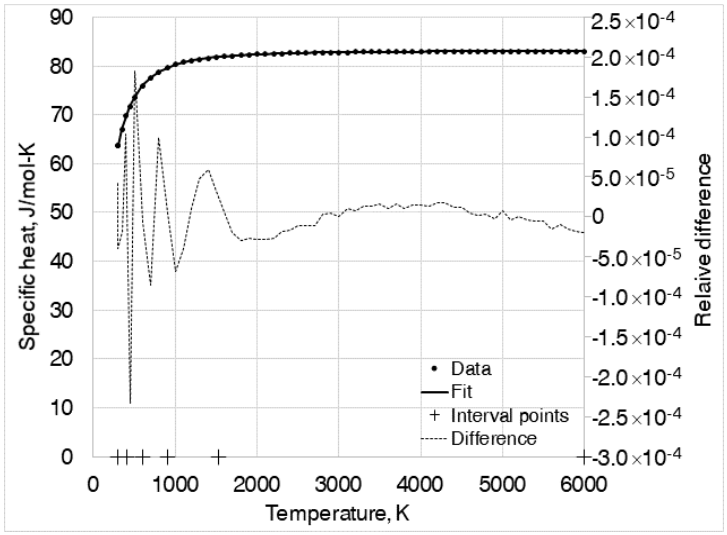


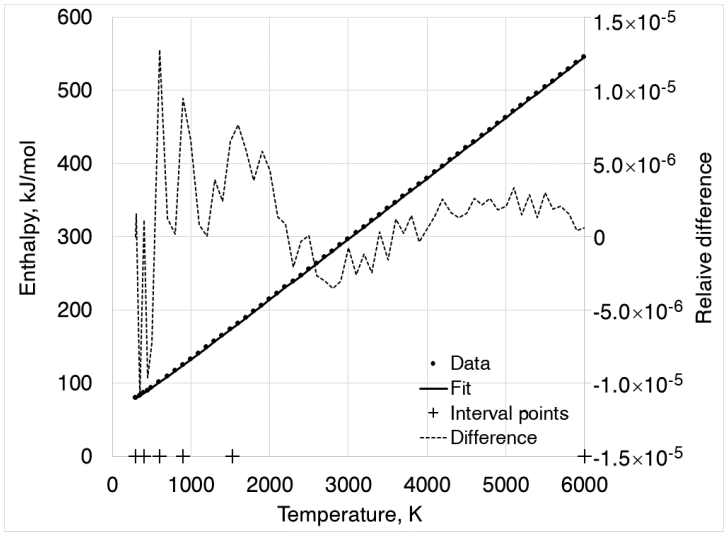


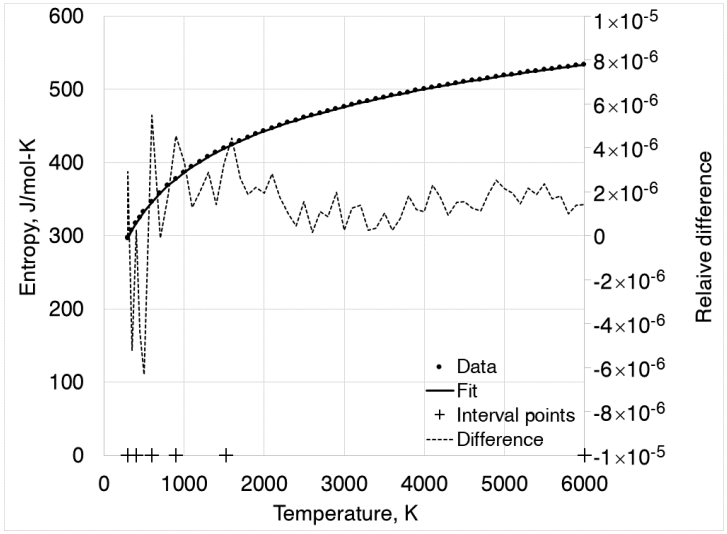


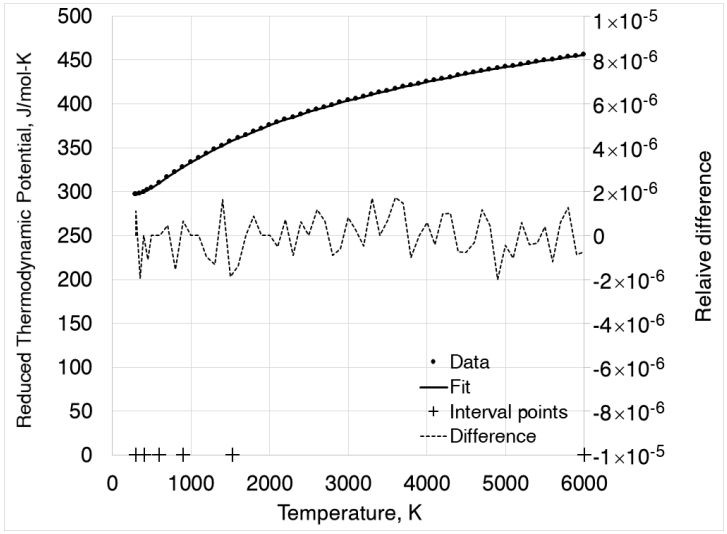
Trichloromethyl (CCl_3) – C1Cl3(g)

MW = 118.37 g/mol – $\Delta_f H^o$(298.15 K) = 79.496 kJ/mol – S^o(298.15 K) = 296.833 J/mol-K – p^o = 0.1 MPa



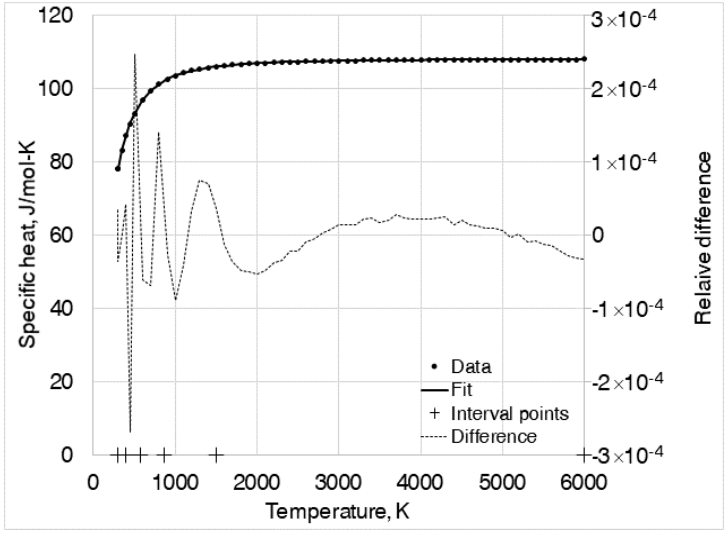


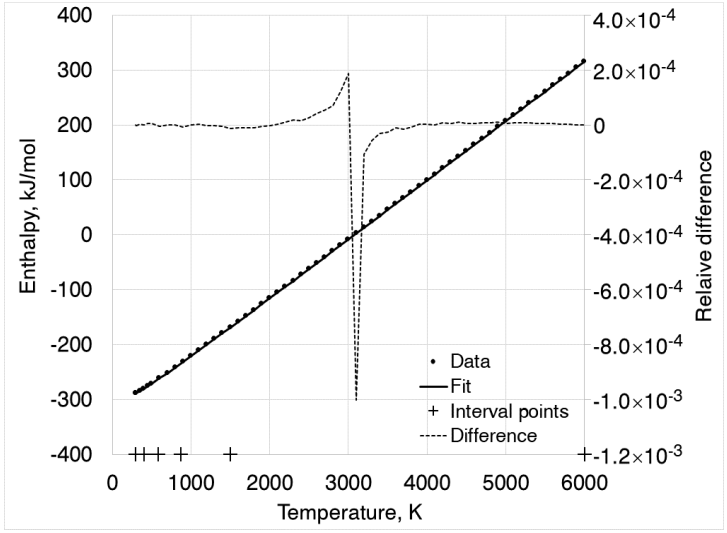


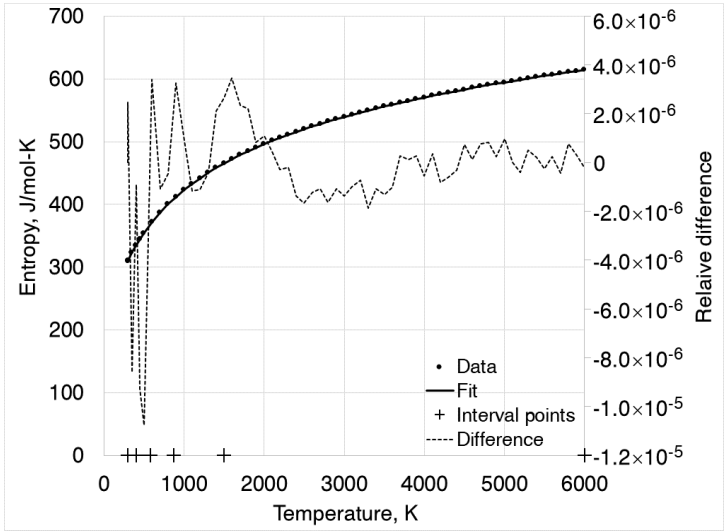


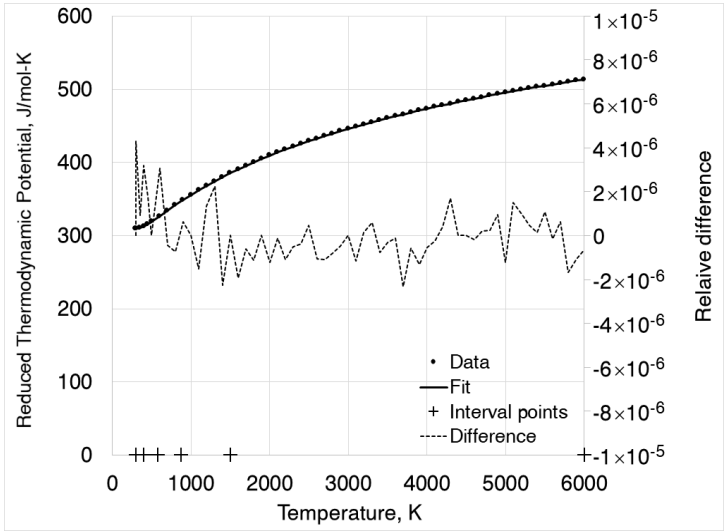
Trichlorofluormethane (CCl_3F) – C1Cl3F1(g)

MW = 137.368 g/mol – $\Delta_f H^o$(298.15 K) = -288.696 kJ/mol – S^o(298.15 K) = 309.735 J/mol-K – p^o = 0.1 MPa



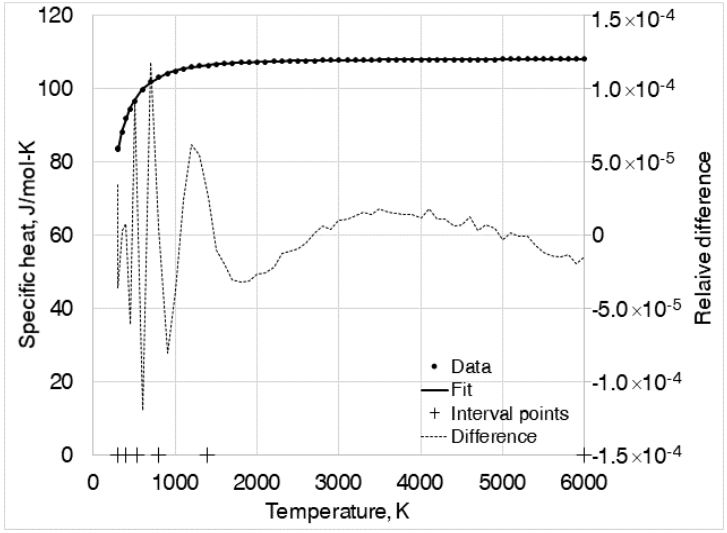


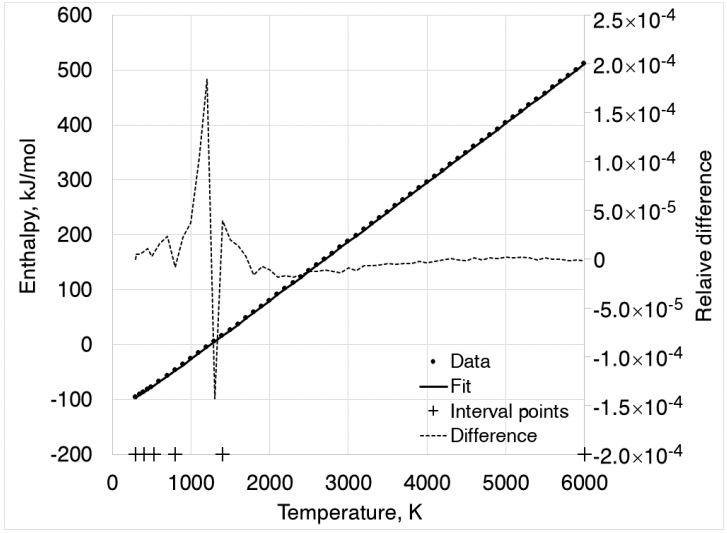


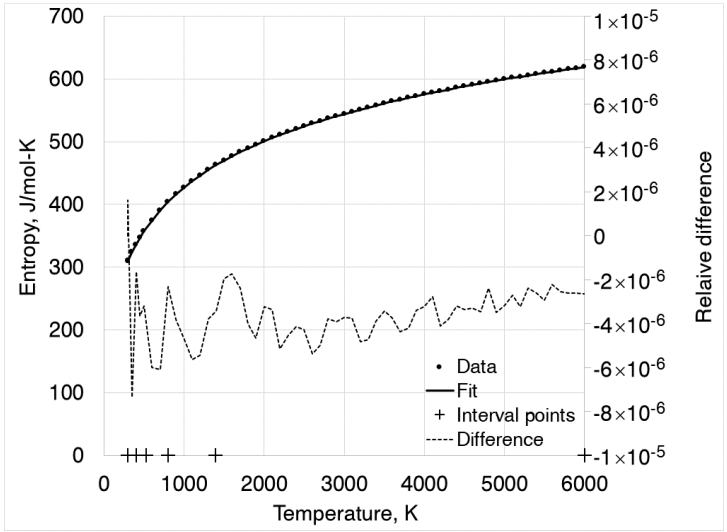


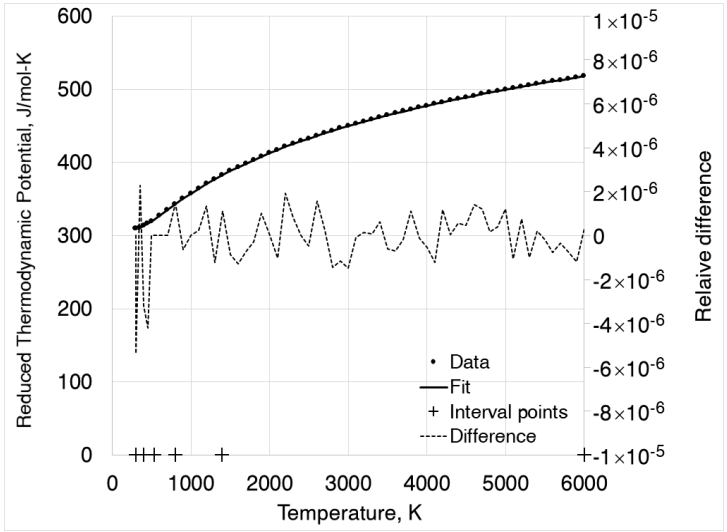
Tetrachloromethane (CCl_4) – C1Cl4(g)

MW = 153.823 g/mol – $\Delta_f H^o$(298.15 K) = -95.981 kJ/mol – S^o(298.15 K) = 309.809 J/mol-K – p^o = 0.1 MPa



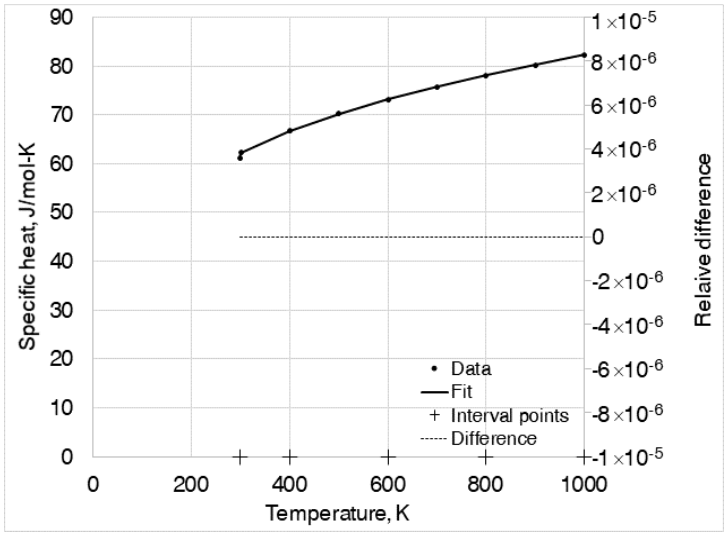


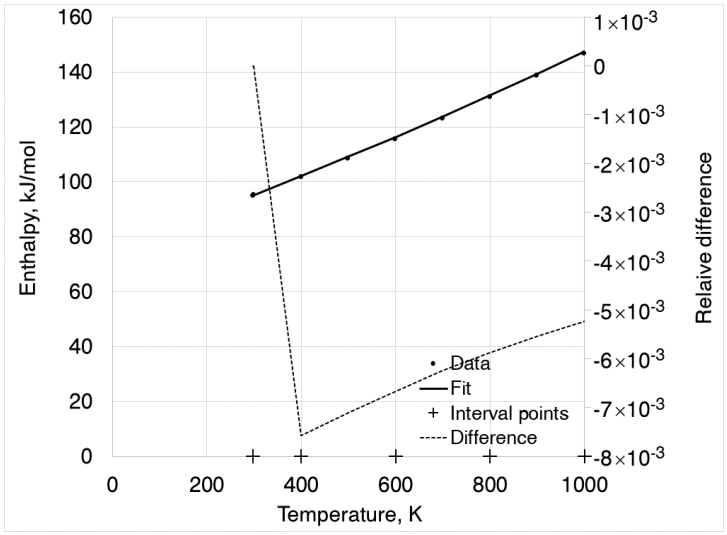


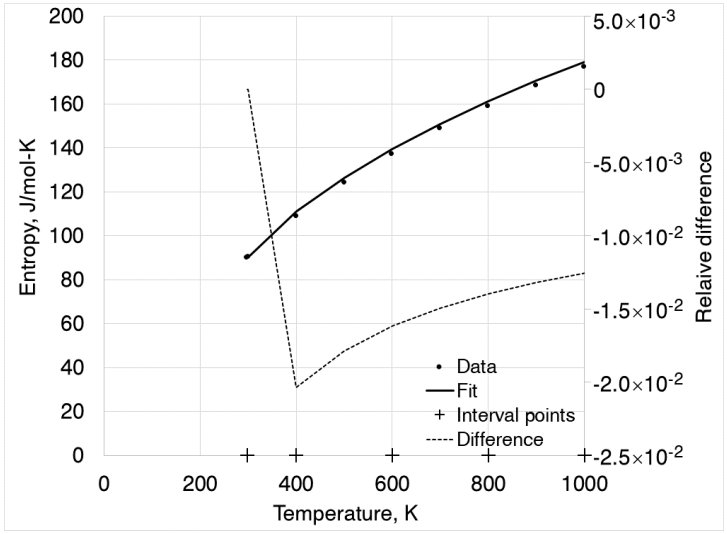


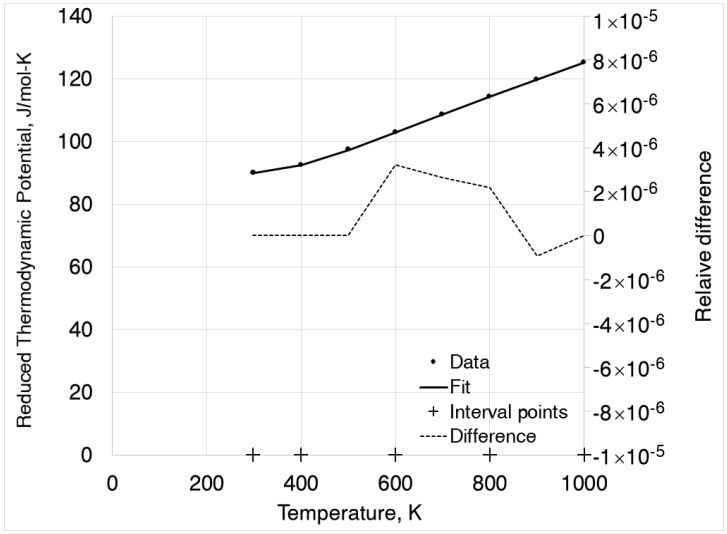
Copper Cyanide (CuCN) – C1Cu1N1(cr)

MW = 89.564 g/mol – $\Delta_f H^o$(298.15 K) = 94.977 kJ/mol – S^o(298.15 K) = 90.002 J/mol-K – p^o = 0.1 MPa



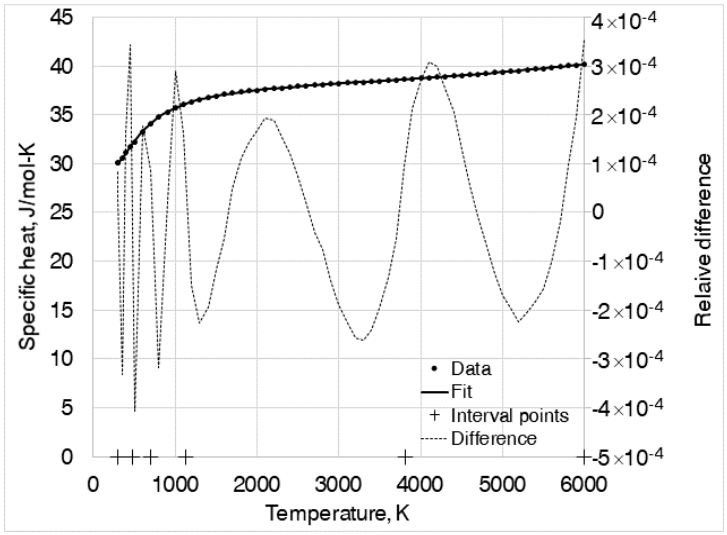


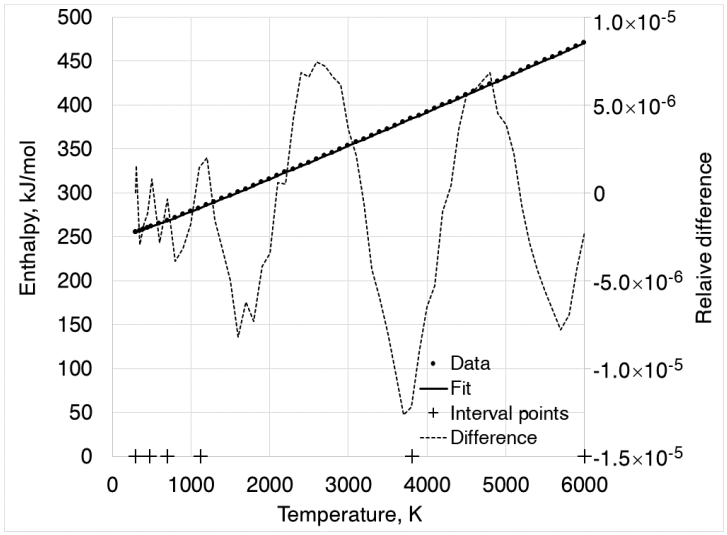


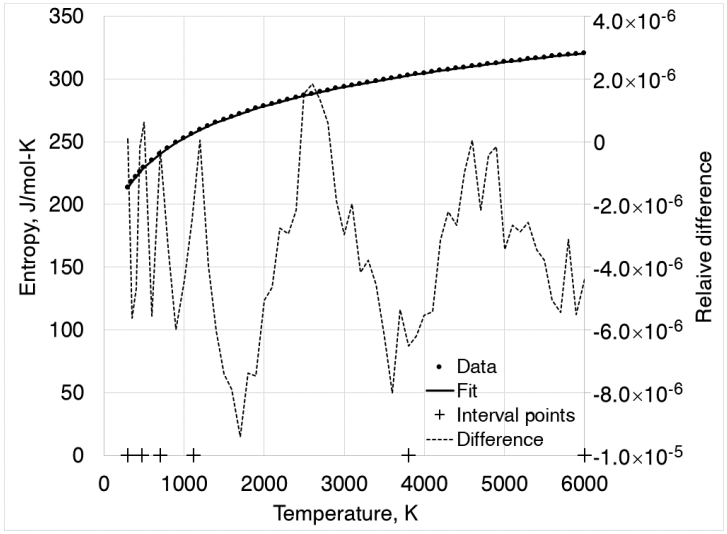


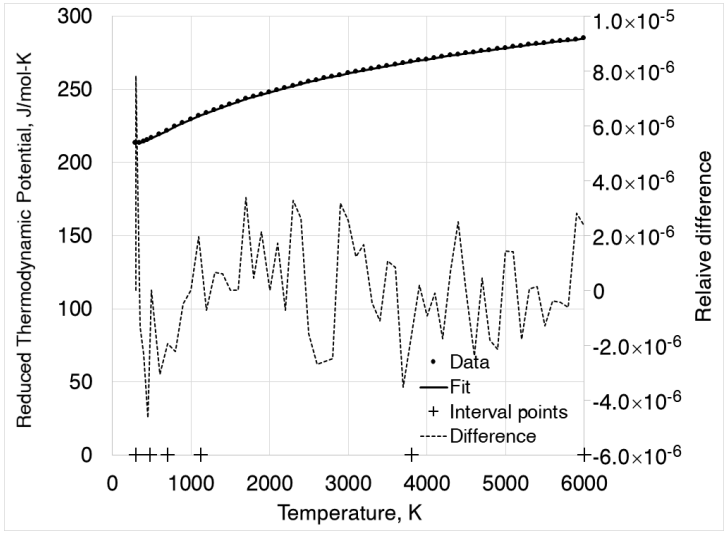
Fluoromethylidyne (CF) – C1F1(g)

MW = 31.009 g/mol – $\Delta_f H^o$(298.15 K) = 255.224 kJ/mol – S^o(298.15 K) = 213.033 J/mol-K – p^o = 0.1 MPa



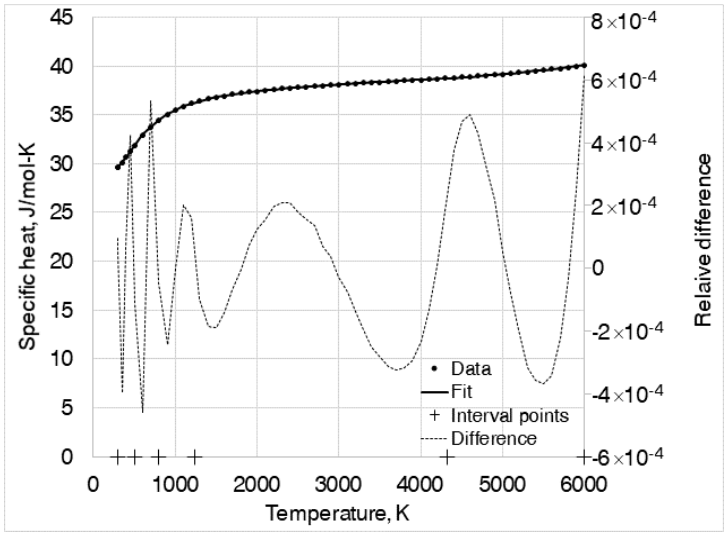


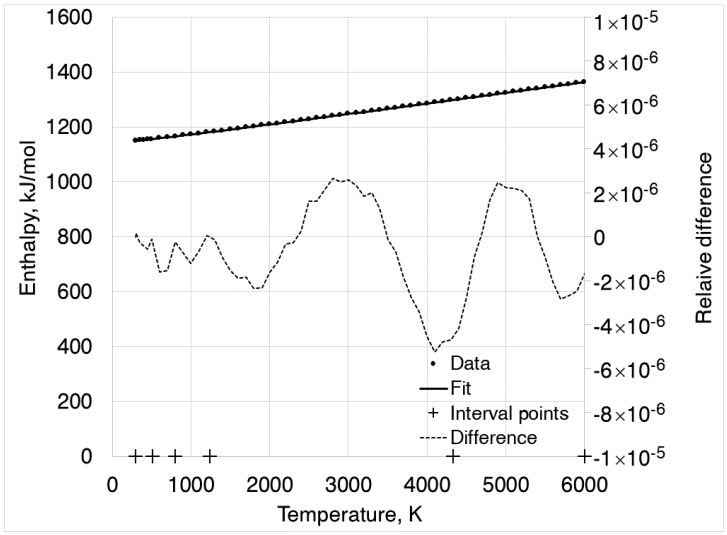


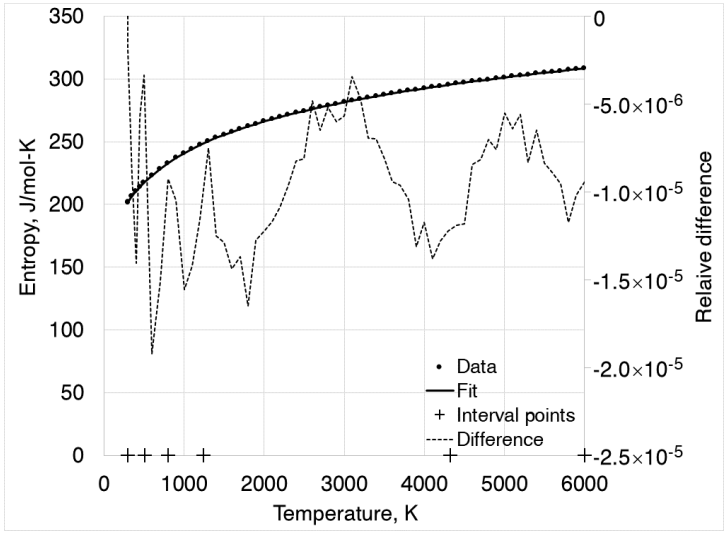


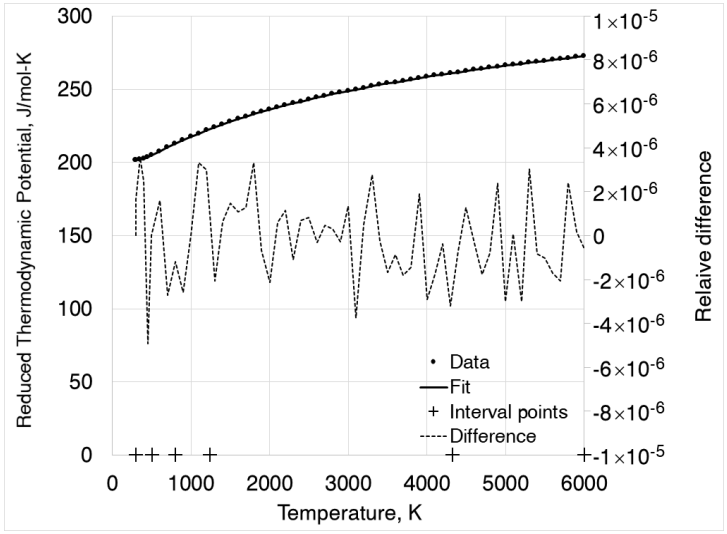
Fluoromethylidyne, Ion (CF^+) – C1F1+(g)

MW = 31.009 g/mol – $\Delta_f H^o$(298.15 K) = 1149.416 kJ/mol – S^o(298.15 K) = 201.507 J/mol-K – p^o = 0.1 MPa



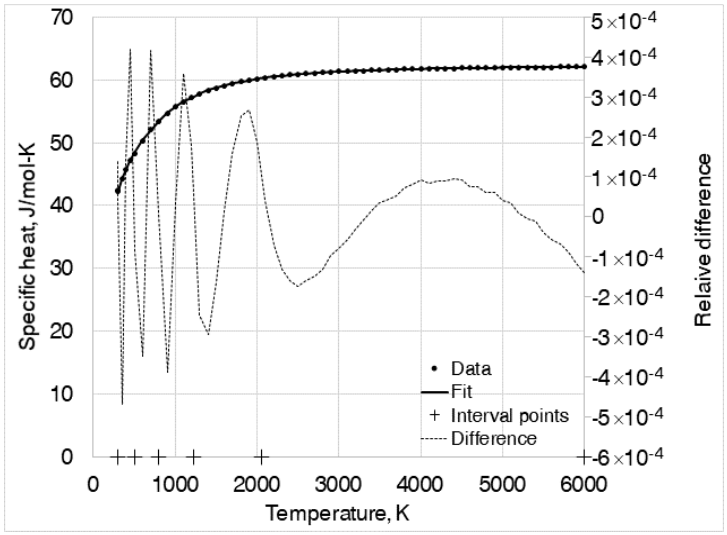


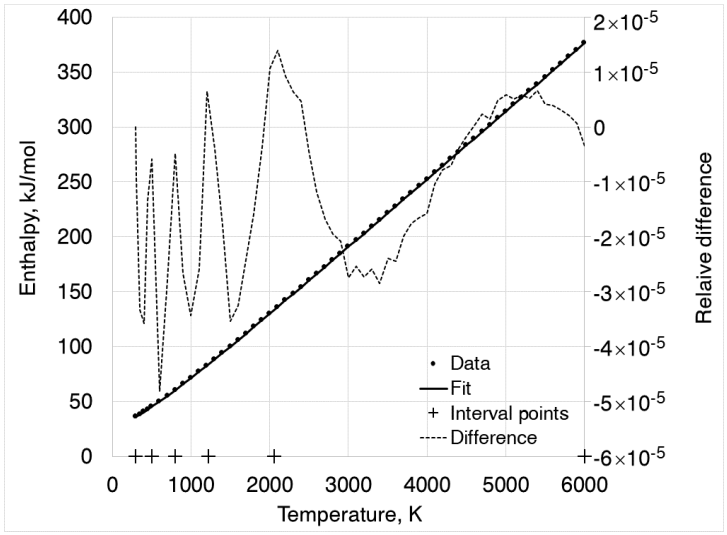


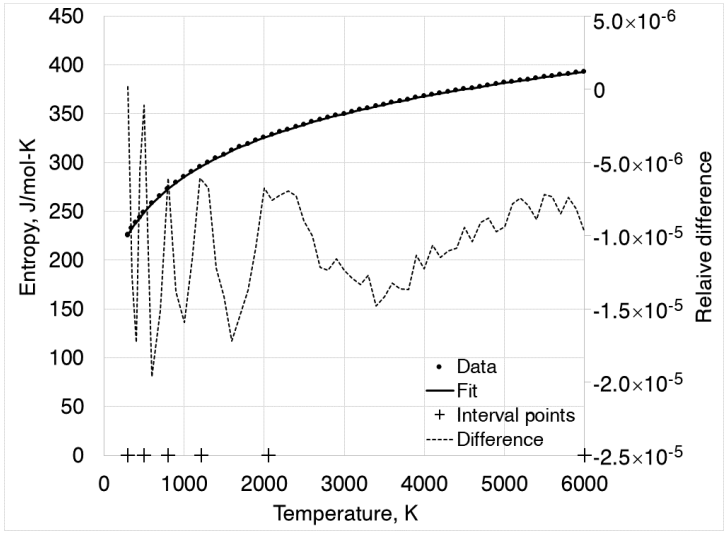


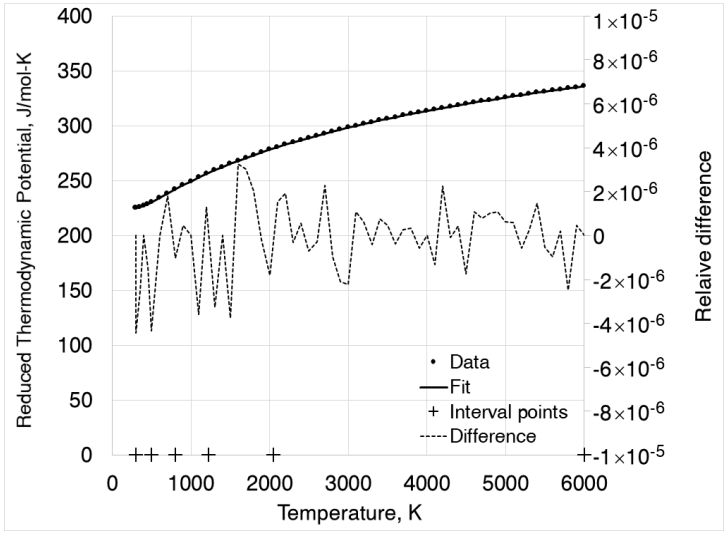
Cyanogen Fluoride (FCN) – C1F1N1(g)

MW = 45.016 g/mol – $\Delta_f H^o$(298.15 K) = 35.982 kJ/mol – S^o(298.15 K) = 225.403 J/mol-K – p^o = 0.1 MPa



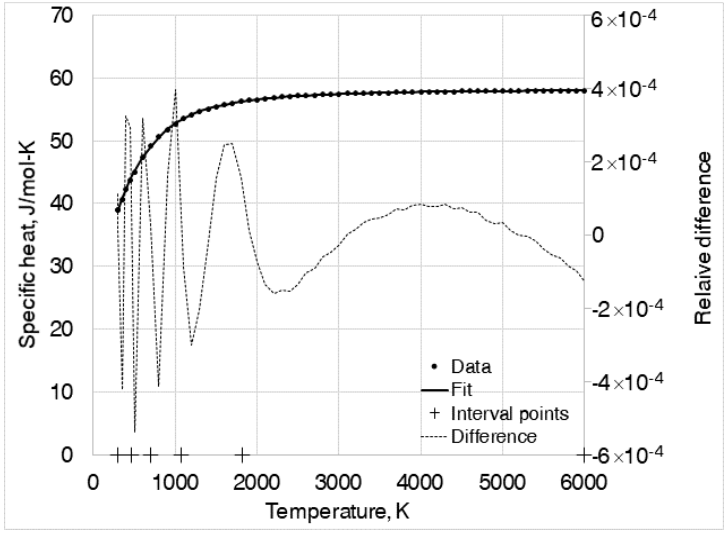


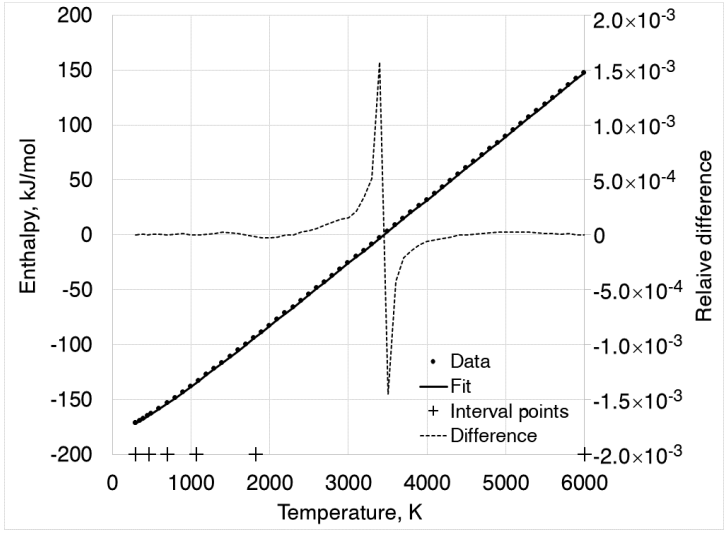


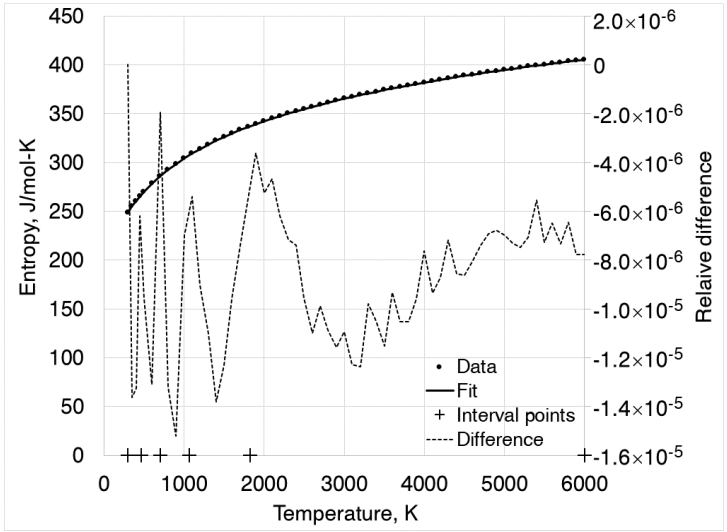


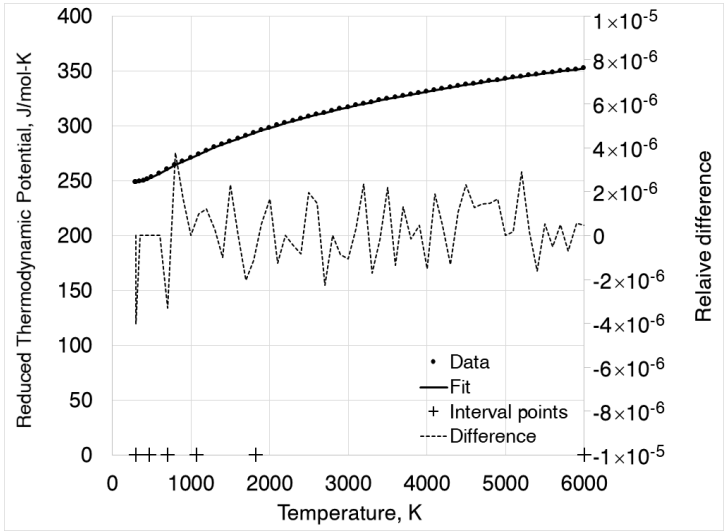
Carbonyl Fluoride (COF) – C1F1O1(g)

MW = 47.009 g/mol – $\Delta_f H^o$(298.15 K) = -171.544 kJ/mol – S^o(298.15 K) = 248.479 J/mol-K – p^o = 0.1 MPa



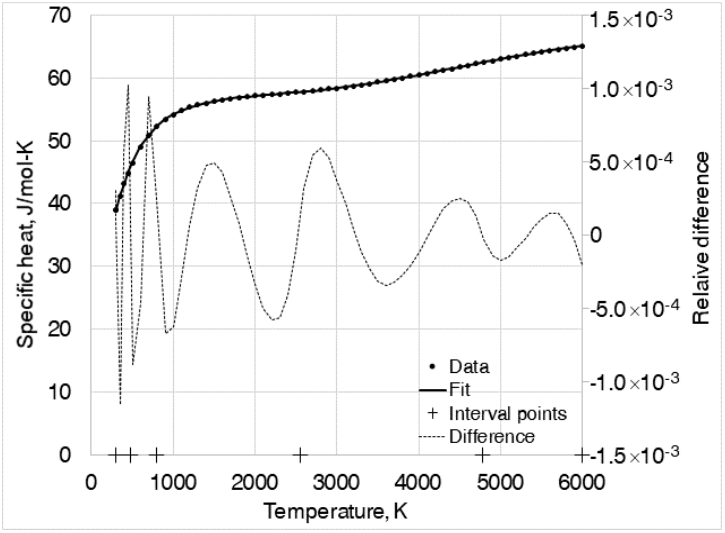


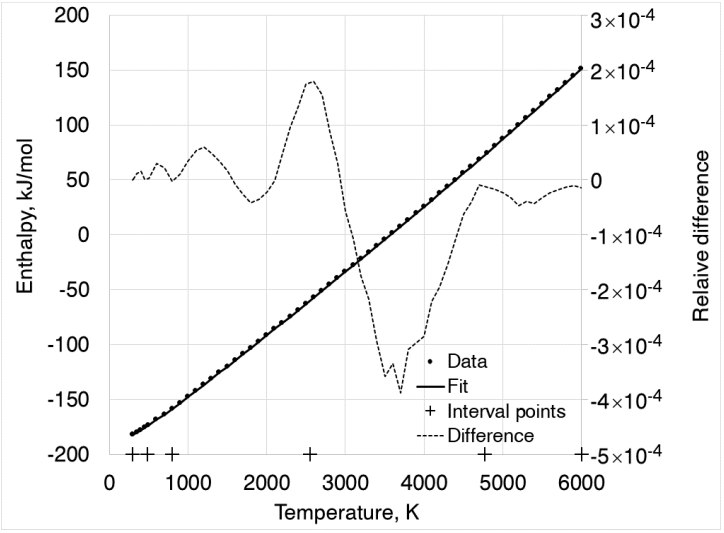


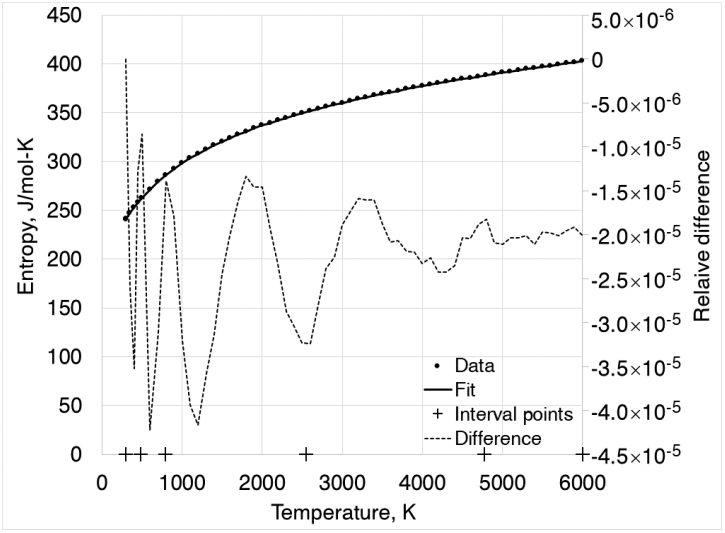


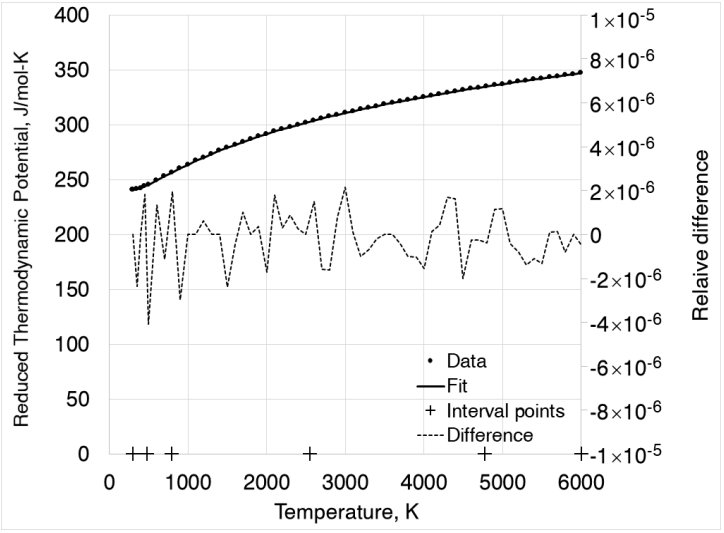
Difluoromethylene (CF_2) – C1F2(g)

MW = 50.008 g/mol – $\Delta_f H^o$(298.15 K) = -182.004 kJ/mol – S^o(298.15 K) = 240.833 J/mol-K – p^o = 0.1 MPa



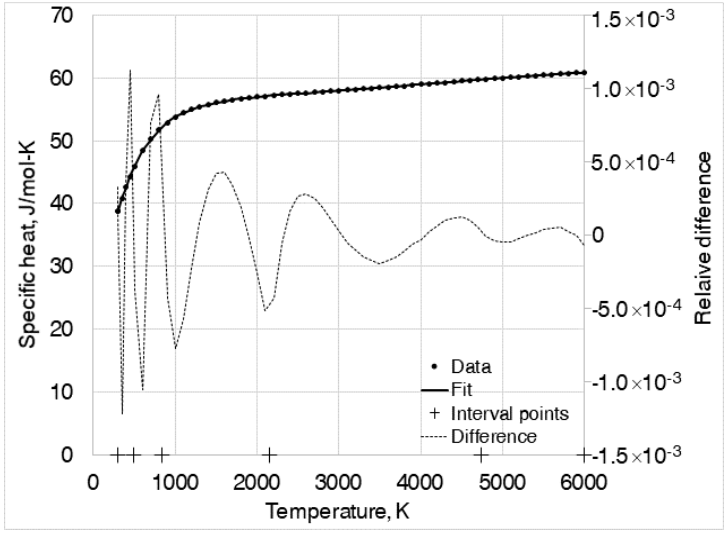


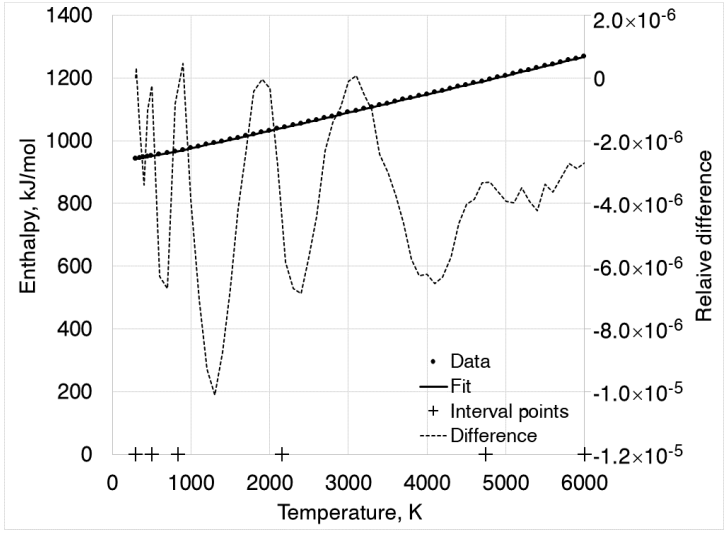


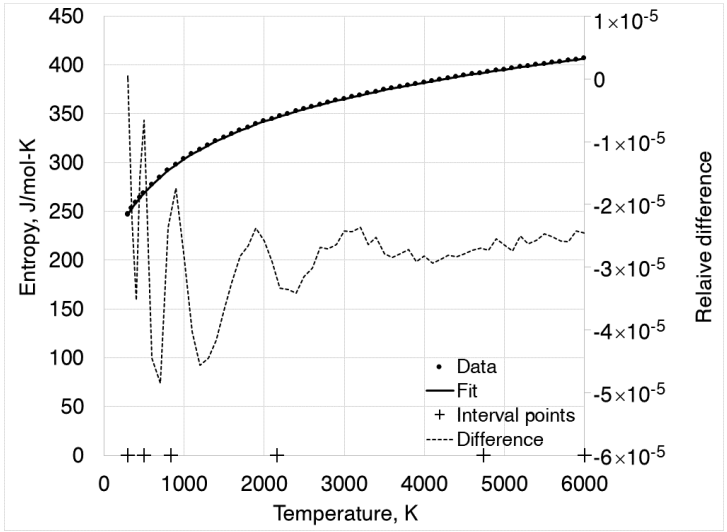


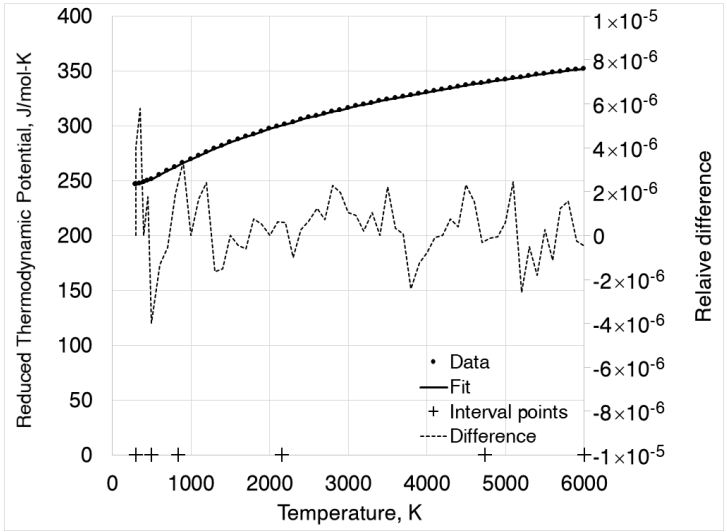
Difluoromethylene, Ion (CF_2^+) – C1F2+(g)

MW = 50.007 g/mol – $\Delta_f H^o$(298.15 K) = 941.793 kJ/mol – S^o(298.15 K) = 246.769 J/mol-K – p^o = 0.1 MPa



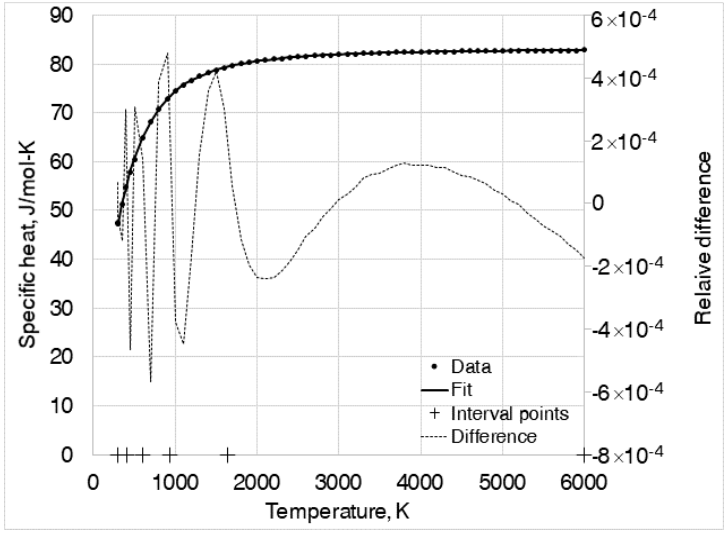


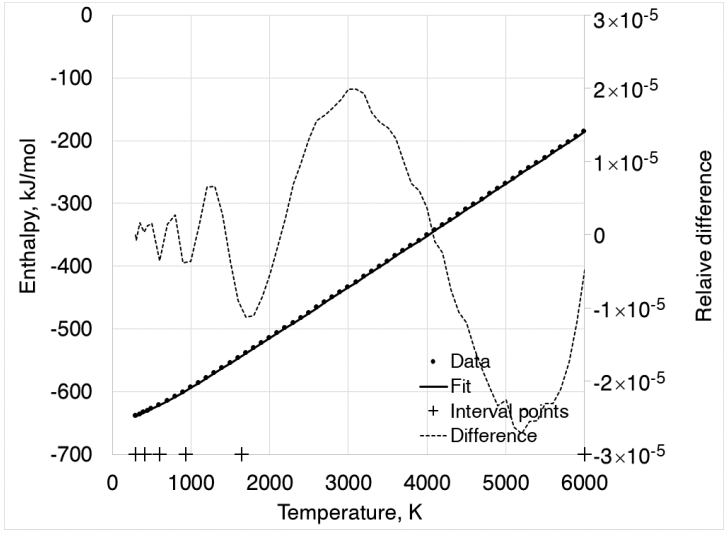


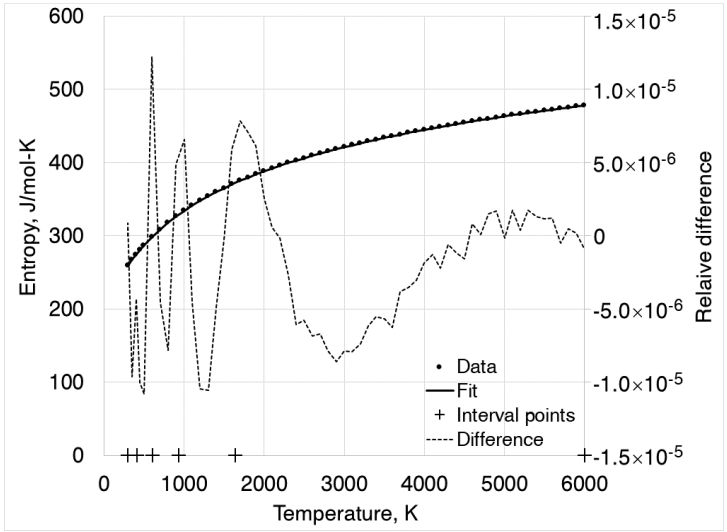


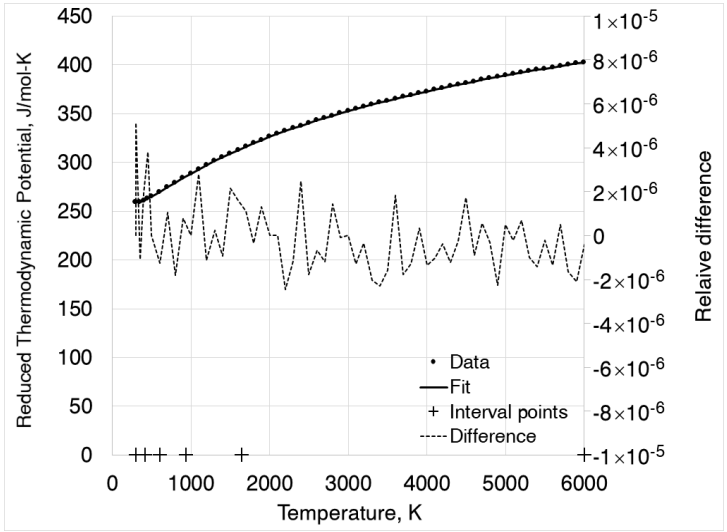
Carbonic Difluoride (COF_2) – C1F2O1(g)

MW = 66.007 g/mol – $\Delta_f H^o$(298.15 K) = -638.897 kJ/mol – S^o(298.15 K) = 258.887 J/mol-K – p^o = 0.1 MPa



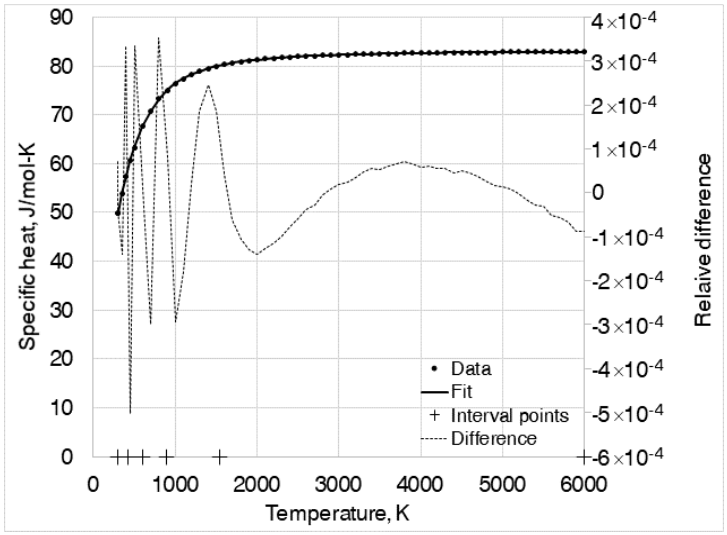


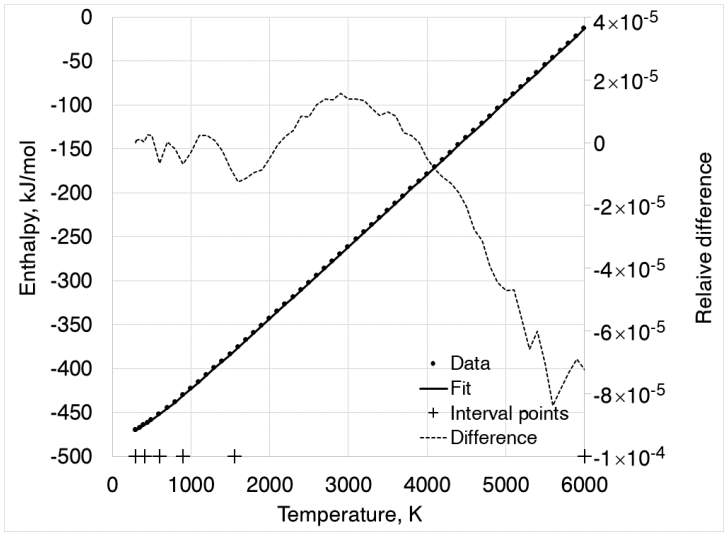


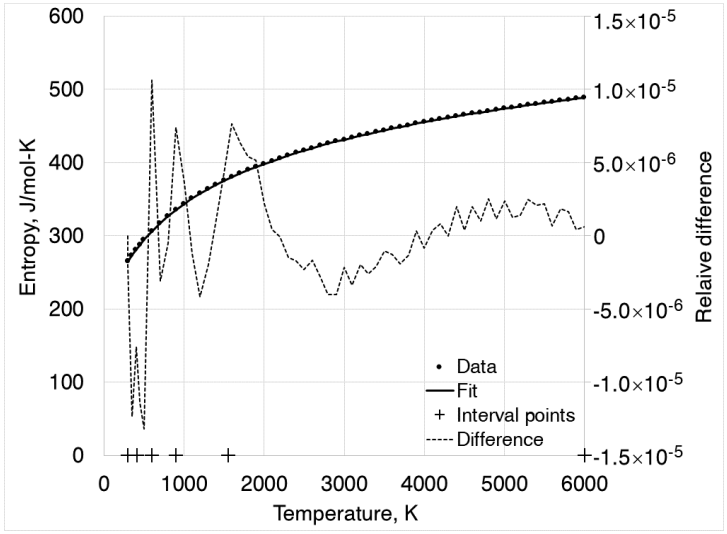


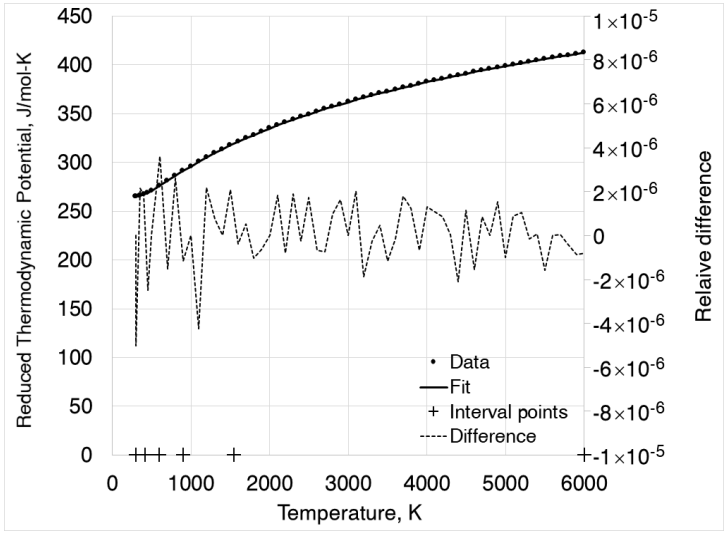
Trifluoromethyl (CF_3) – C1F3(g)

MW = 69.006 g/mol – $\Delta_f H^o$(298.15 K) = -470.282 kJ/mol – S^o(298.15 K) = 265.082 J/mol-K – p^o = 0.1 MPa



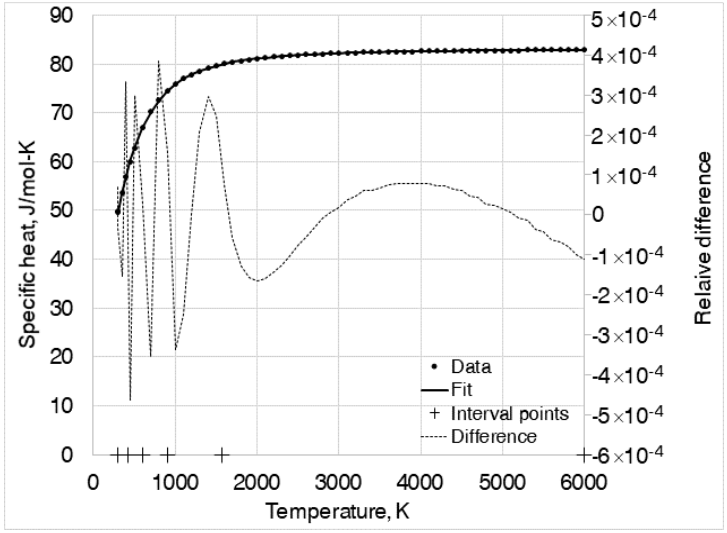


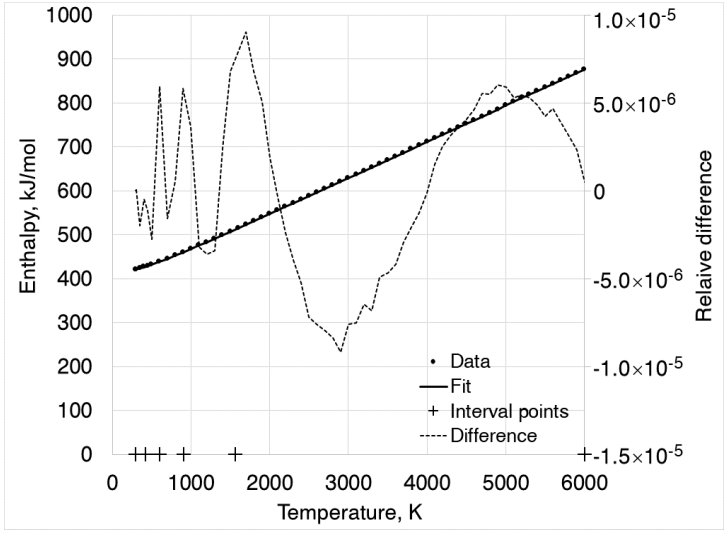


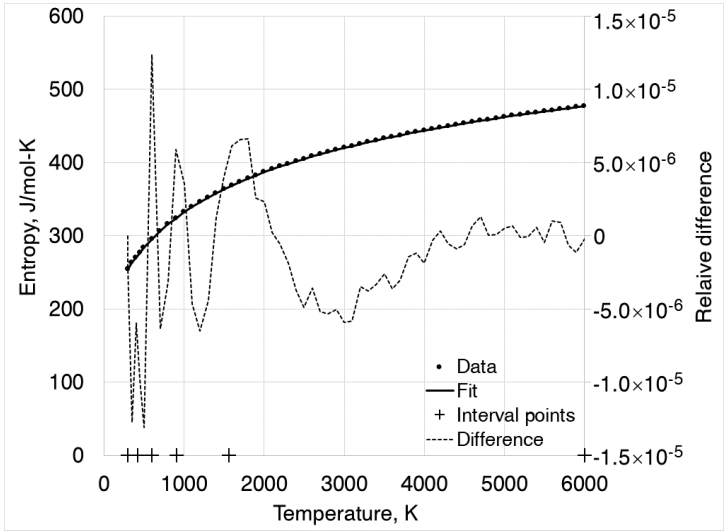


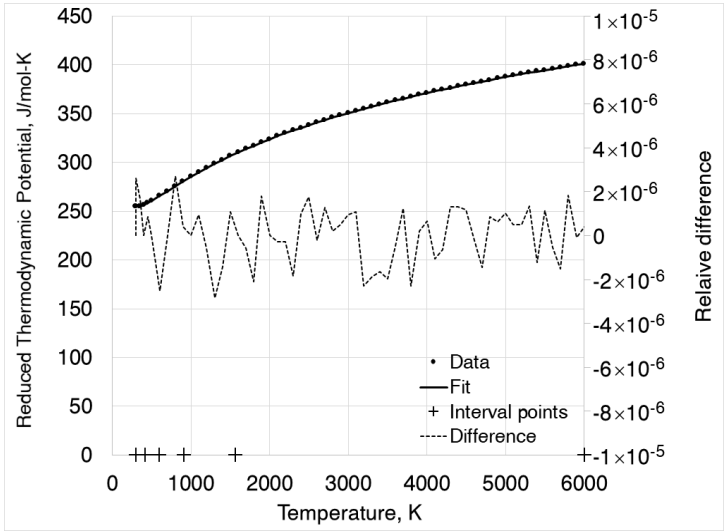
Trifluoromethyl, Ion (CF_3^+) – C1F3+(g)

MW = 69.006 g/mol – $\Delta_f H^o$(298.15 K) = 421.007 kJ/mol – S^o(298.15 K) = 254.532 J/mol-K – p^o = 0.1 MPa



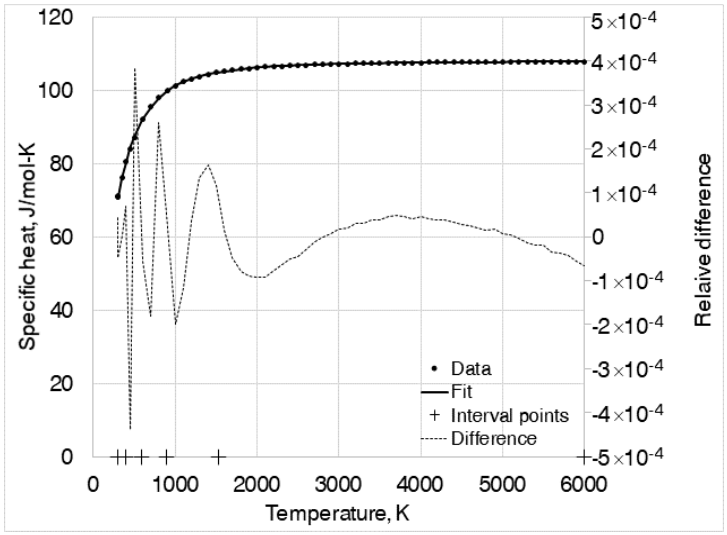


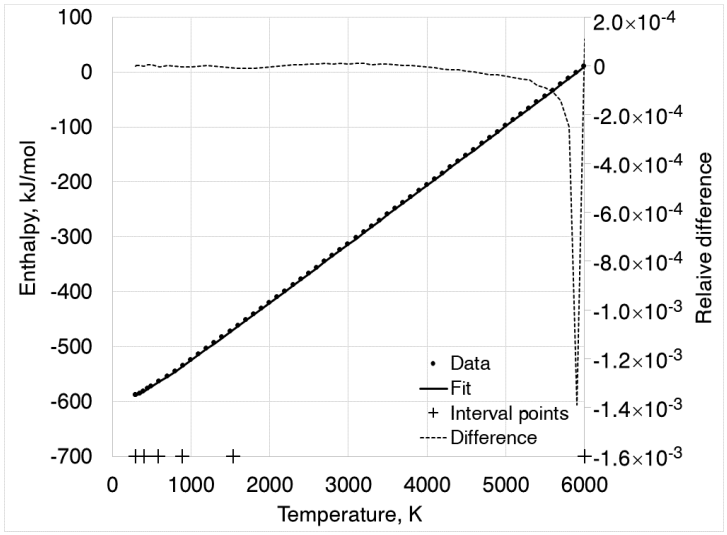


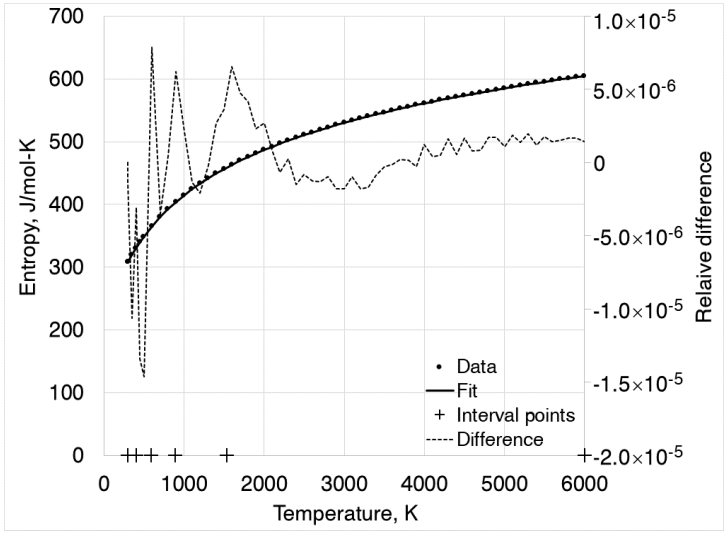


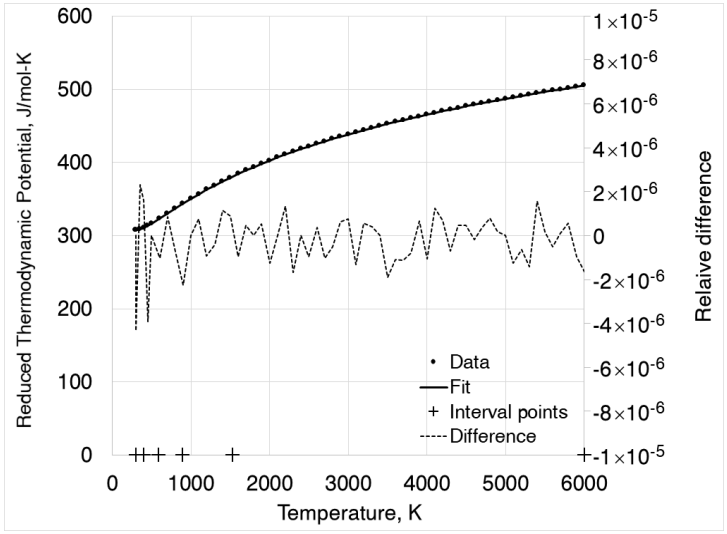
Trifluoriodomethane (CIF_3) – C1F3I1(g)

MW = 195.911 g/mol – $\Delta_f H^o$(298.15 K) = -589.107 kJ/mol – S^o(298.15 K) = 307.629 J/mol-K – p^o = 0.1 MPa



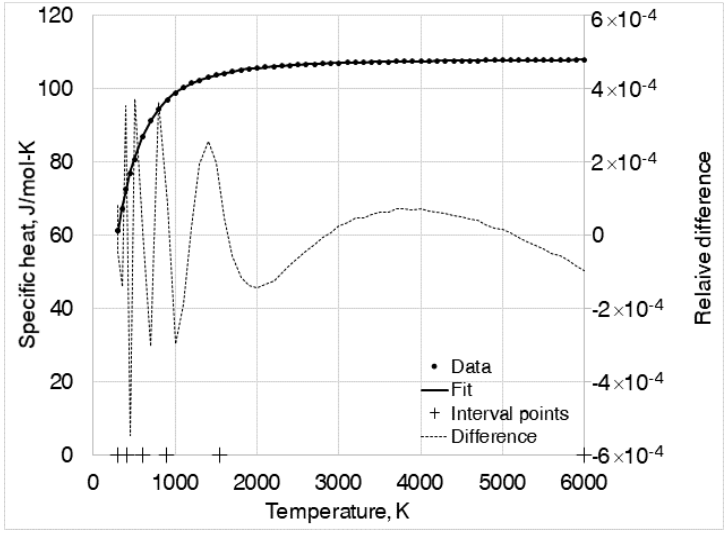


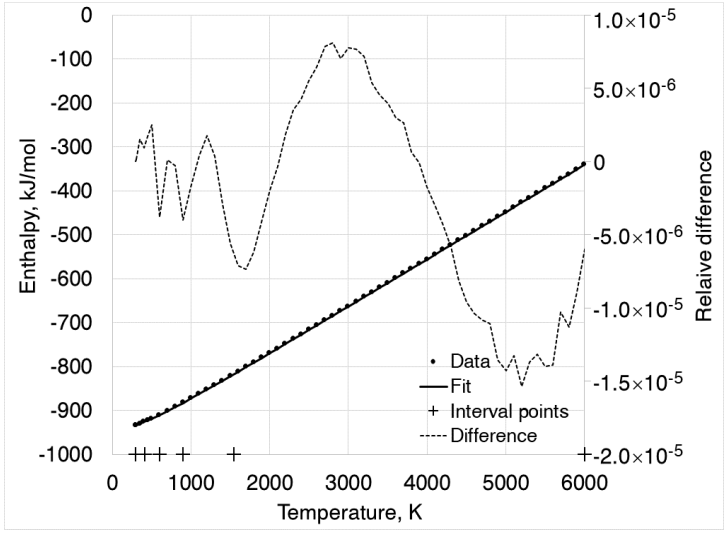


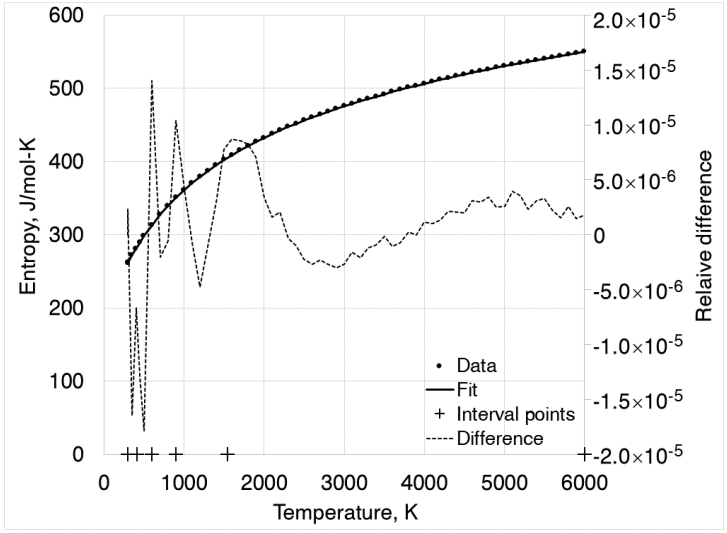


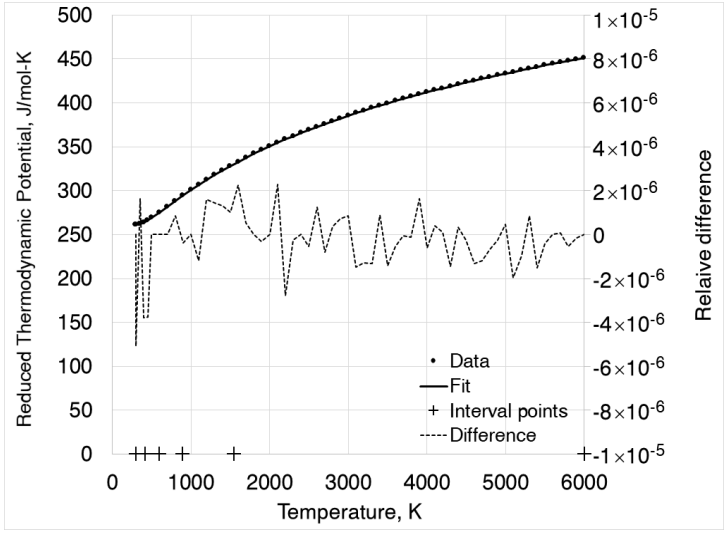
Tetrafluoromethane (CF_4) – C1F4(g)

MW = 88.005 g/mol – $\Delta_f H^o$(298.15 K) = -933.199 kJ/mol – S^o(298.15 K) = 261.419 J/mol-K – p^o = 0.1 MPa



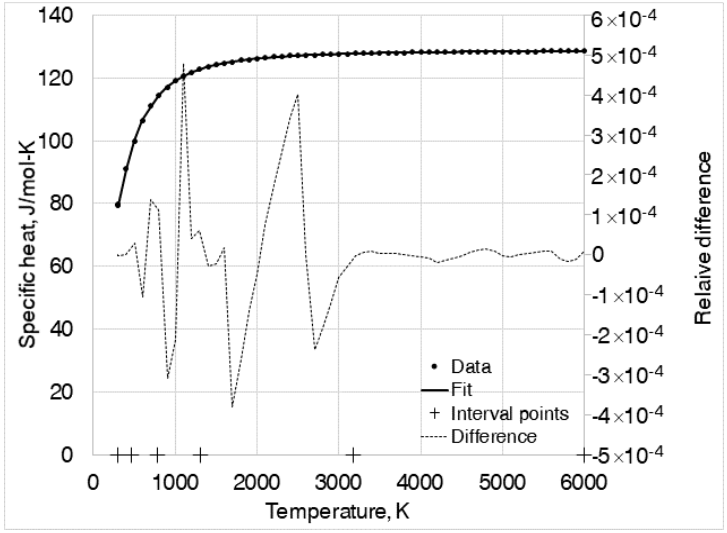


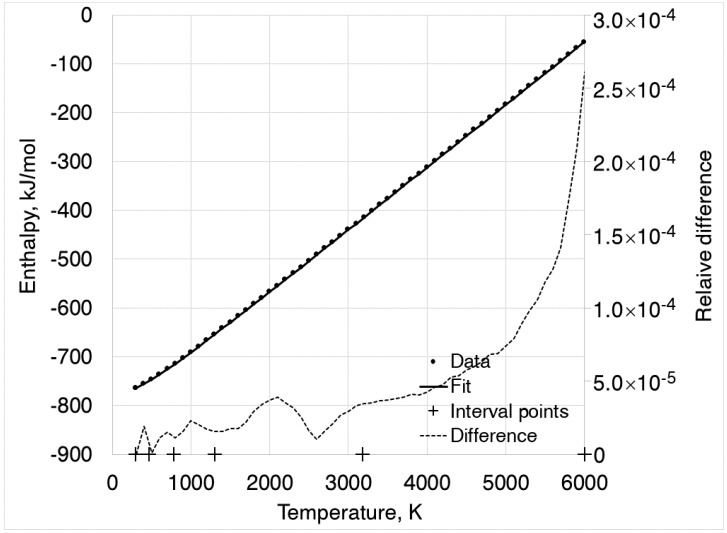


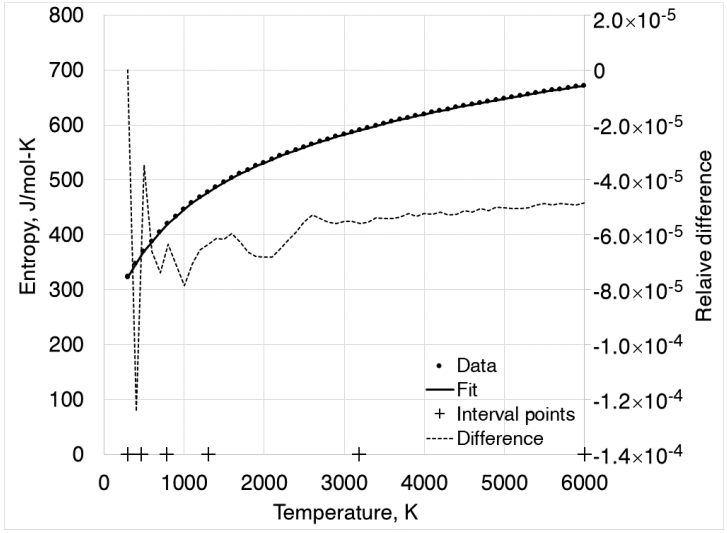


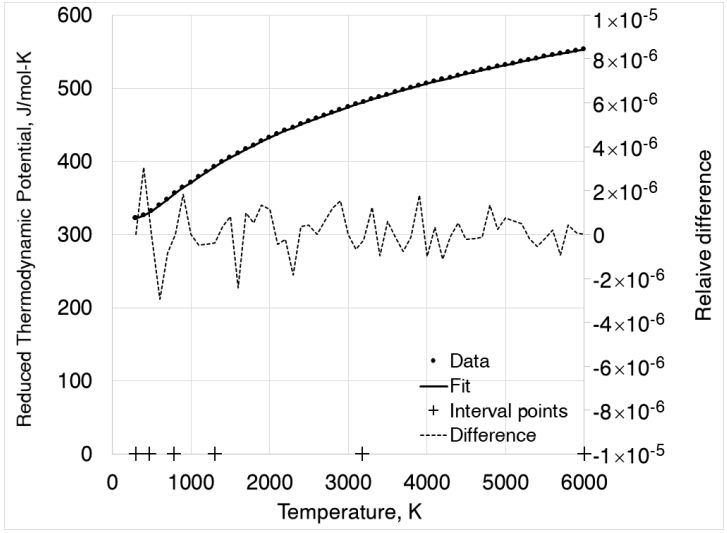
Trifluoromethyl Hypofluorite (CF_3OF) – C1F4O1(g)

MW = 104.004 g/mol – $\Delta_f H^o$(298.15 K) = -764.835 kJ/mol – S^o(298.15 K) = 322.533 J/mol-K – p^o = 0.1 MPa



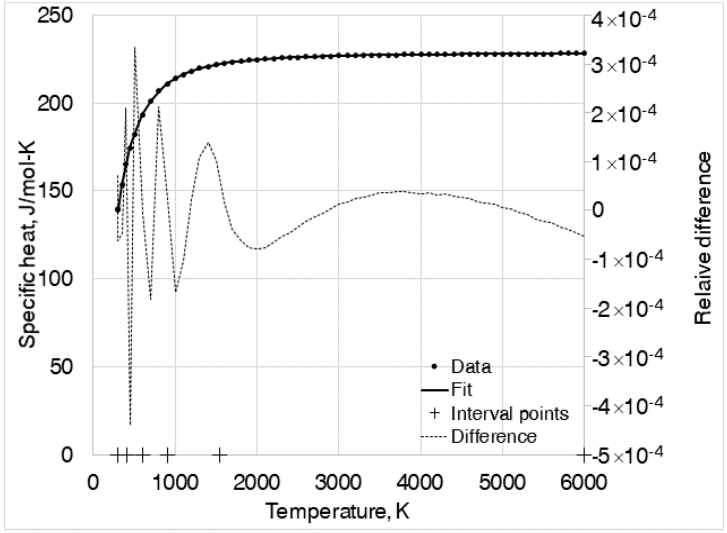


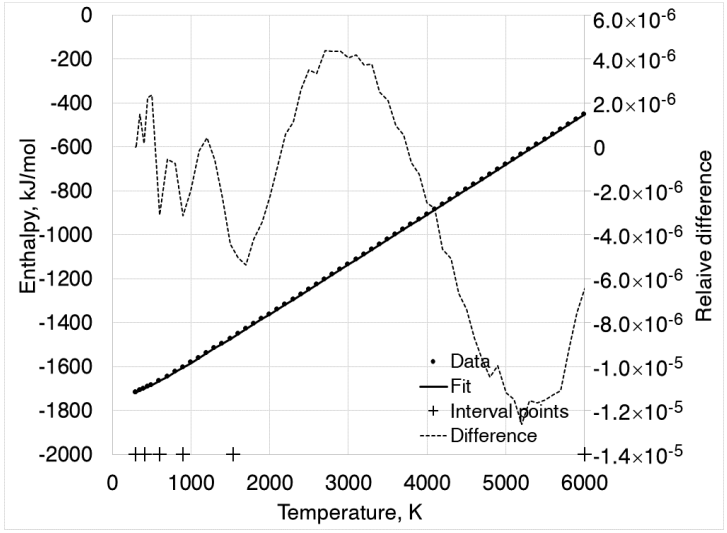


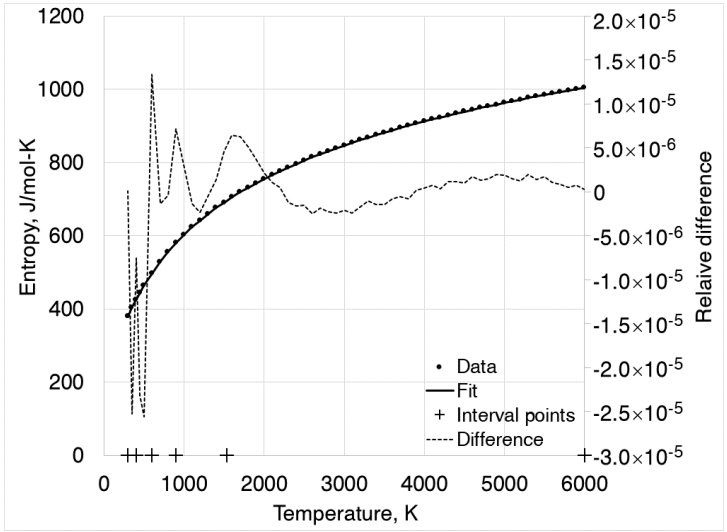


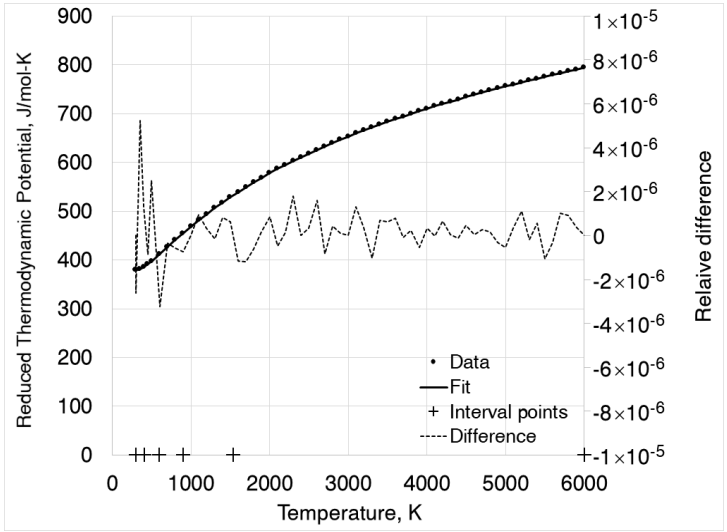
Pentafluoro(Trifluoromethyl)Sulfur (CF_3SF_5) – C1F8S1(g)

MW = 196.06 g/mol – $\Delta_f H^o$(298.15 K) = -1717.051 kJ/mol – S^o(298.15 K) = 379.967 J/mol-K – p^o = 0.1 MPa



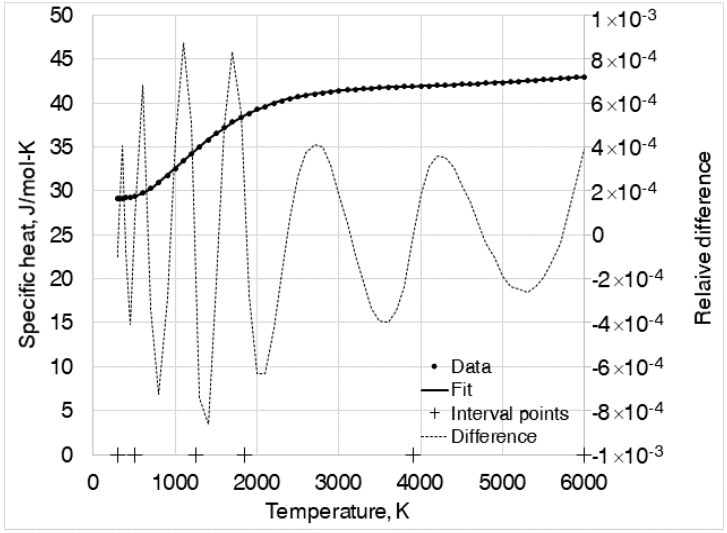


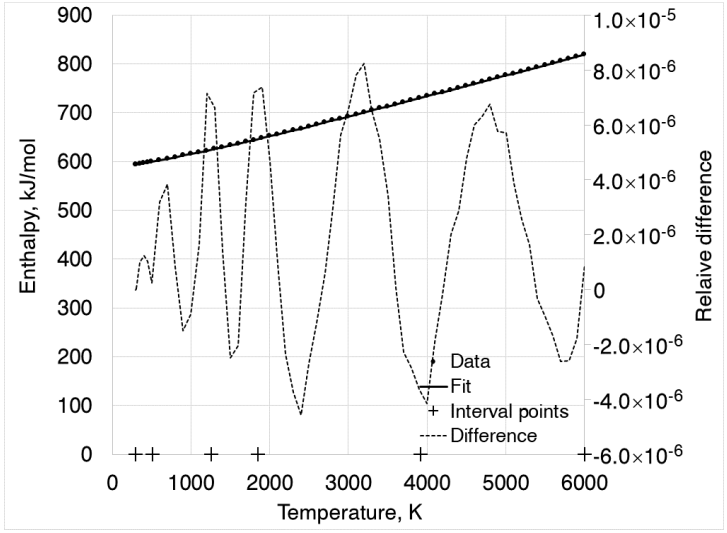


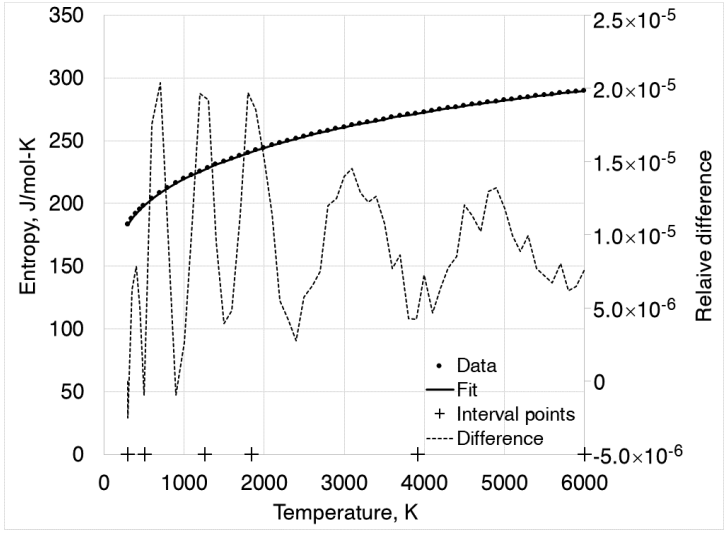


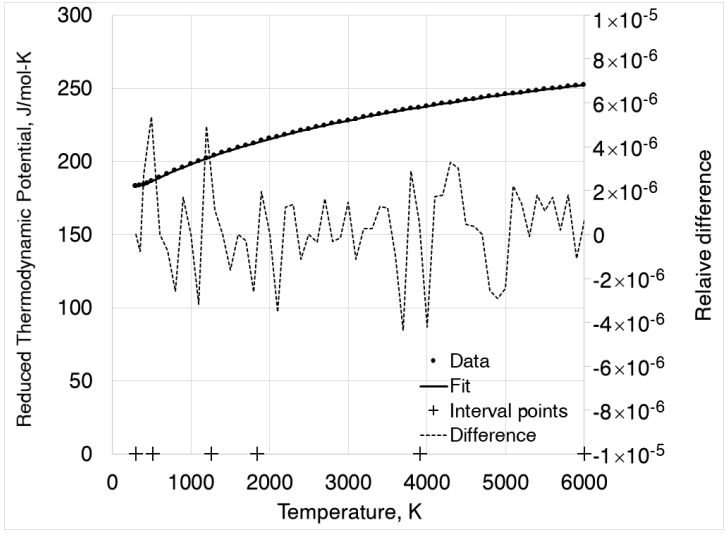
Methylidyne (CH) – C1H1(g)

MW = 13.019 g/mol – $\Delta_f H^o$(298.15 K) = 594.128 kJ/mol – S^o(298.15 K) = 183.04 J/mol-K – p^o = 0.1 MPa



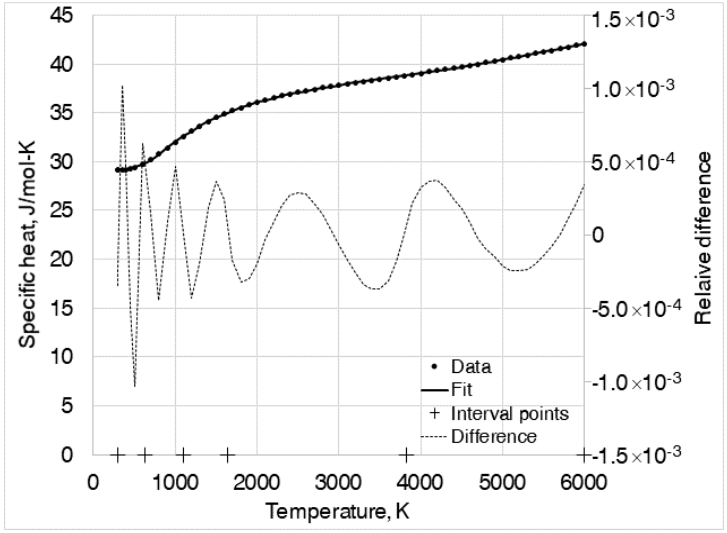


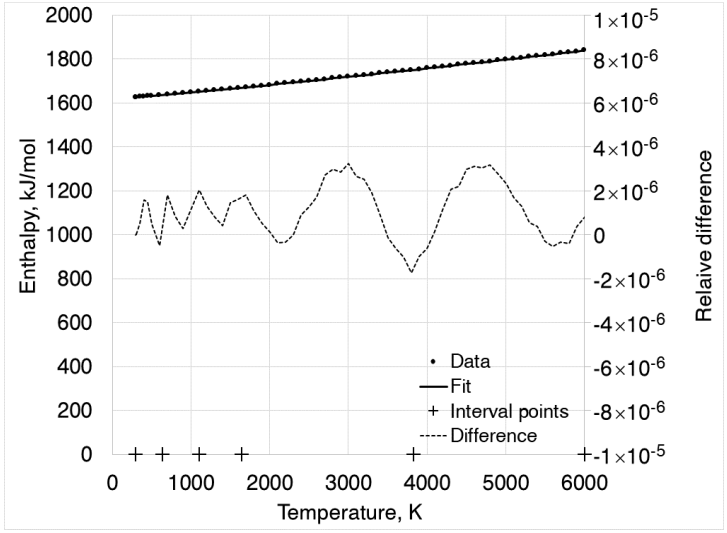


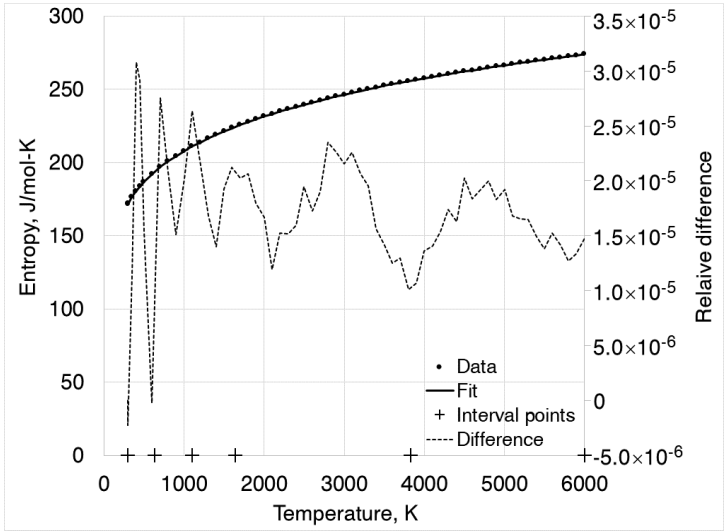


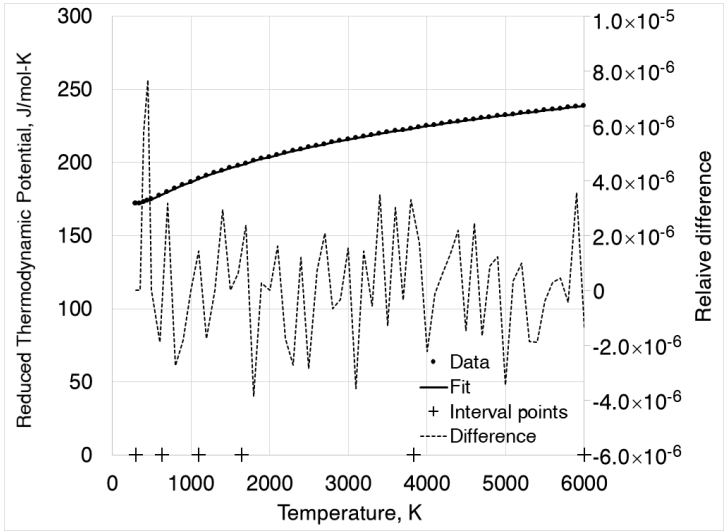
Methylidyne, Ion (CH^+) – C1H1+(g)

MW = 13.018 g/mol – $\Delta_f H^o$(298.15 K) = 1626.739 kJ/mol – S^o(298.15 K) = 171.671 J/mol-K – p^o = 0.1 MPa



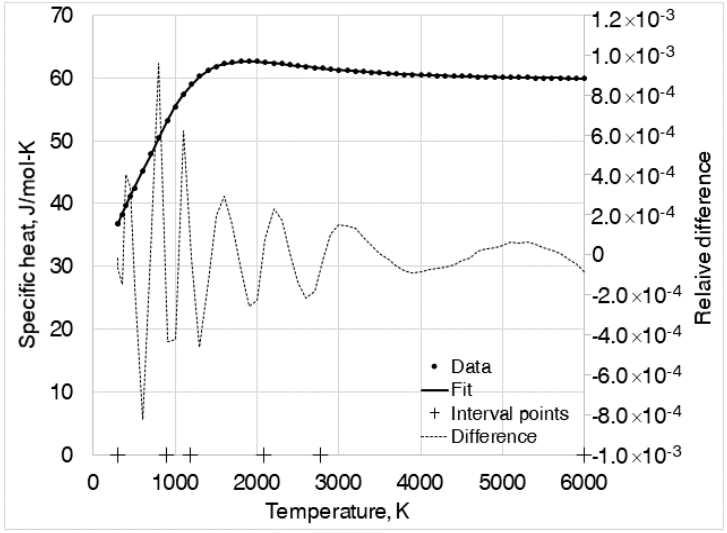


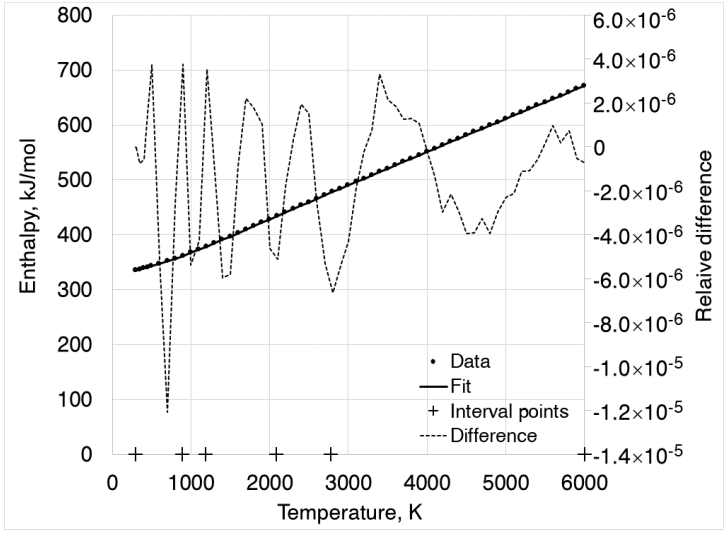


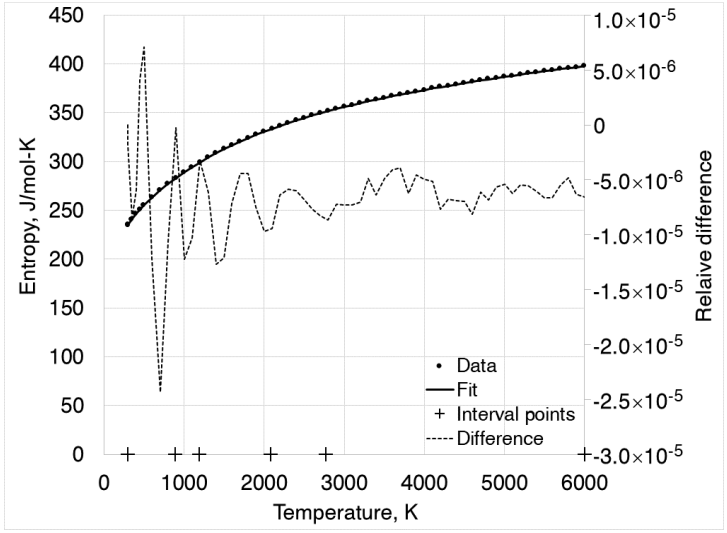


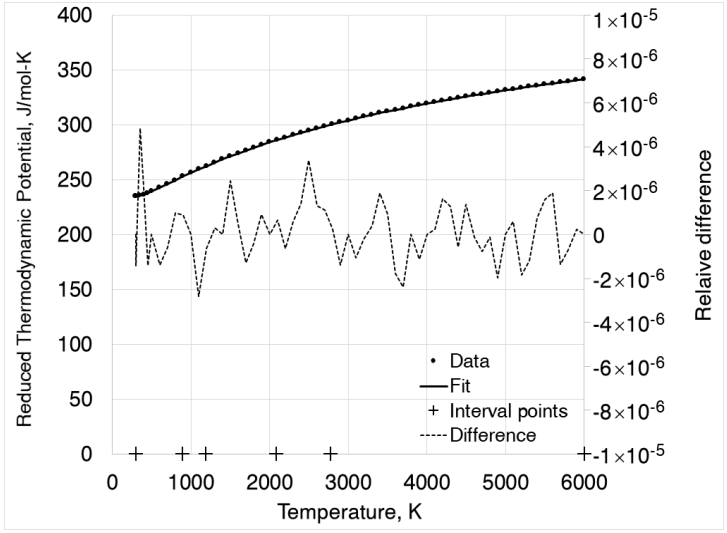
Chloromethylene (CHCl) – C1H1Cl1(g)

MW = 48.472 g/mol – $\Delta_f H^o$(298.15 K) = 334.72 kJ/mol – S^o(298.15 K) = 234.912 J/mol-K – p^o = 0.1 MPa



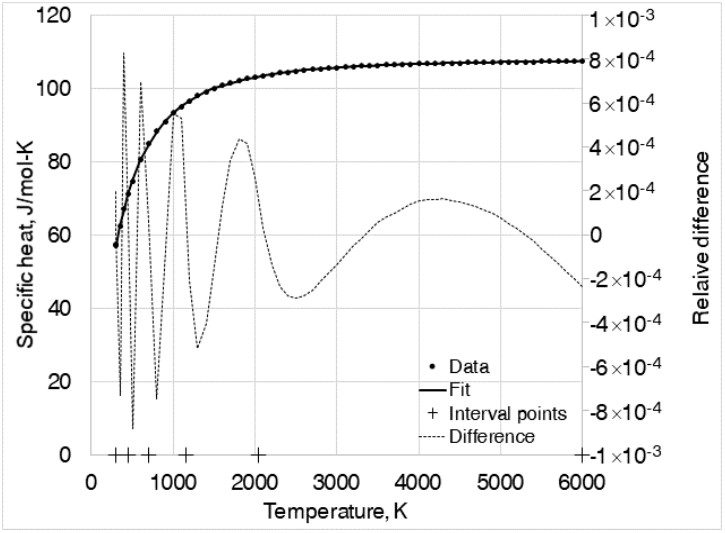


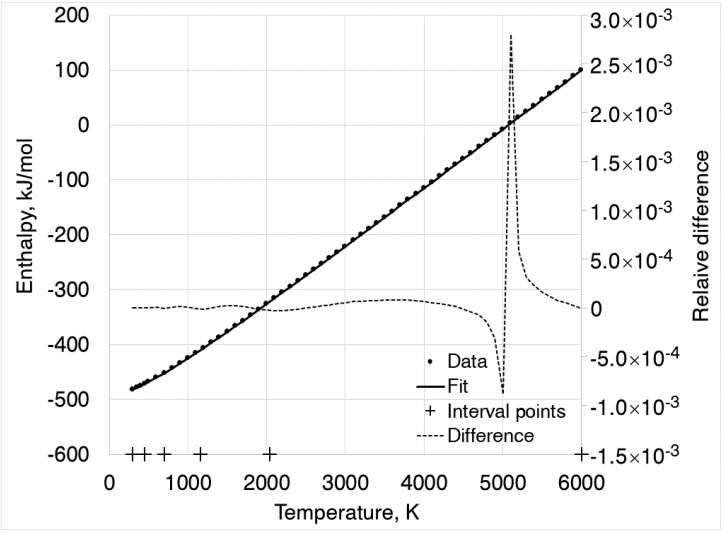


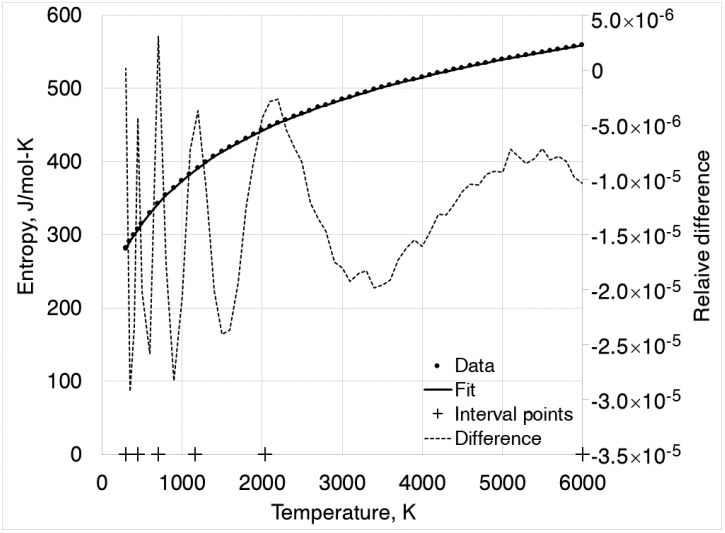


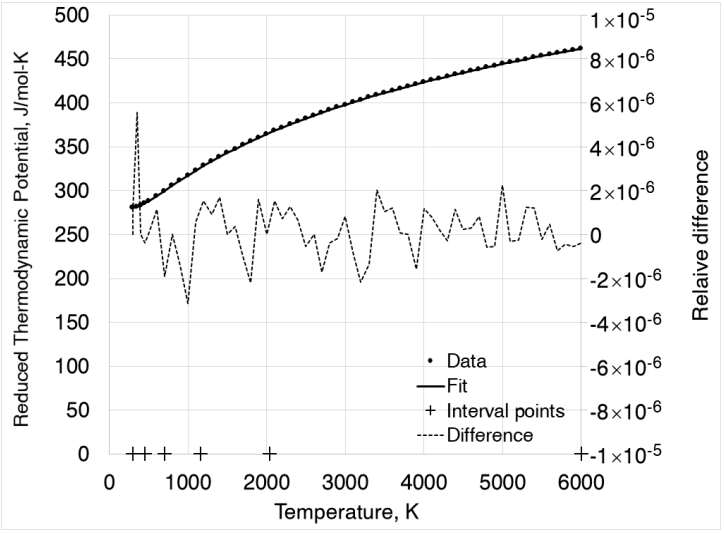
Chlorodifluoromethane ($CHClF_2$) – C1H1Cl1F2(g)

MW = 86.469 g/mol – $\Delta_f H^o$(298.15 K) = -481.578 kJ/mol – S^o(298.15 K) = 280.968 J/mol-K – p^o = 0.1 MPa



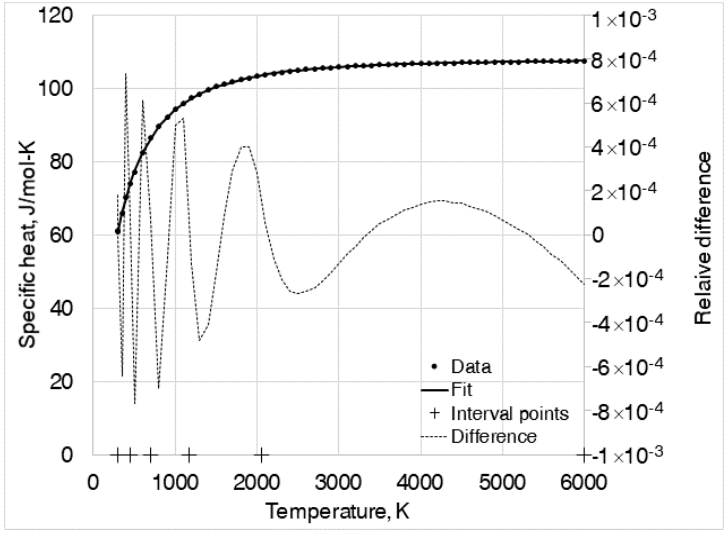


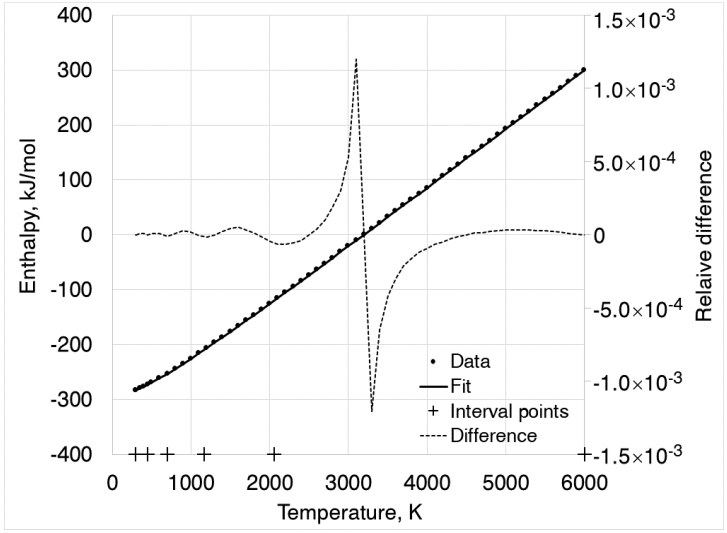


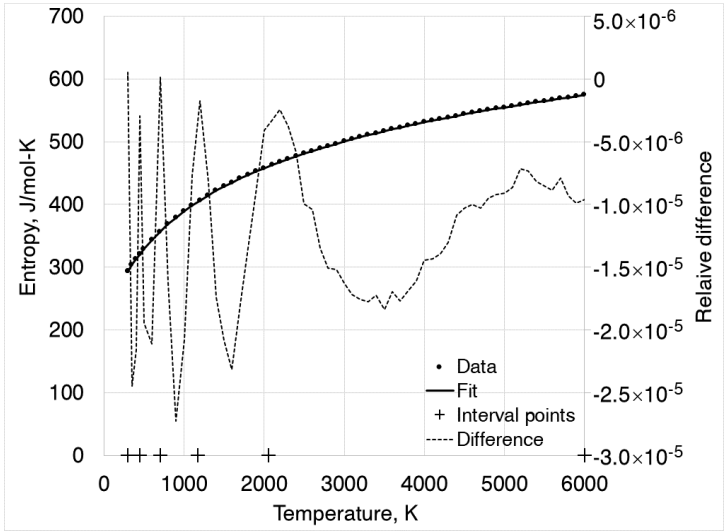


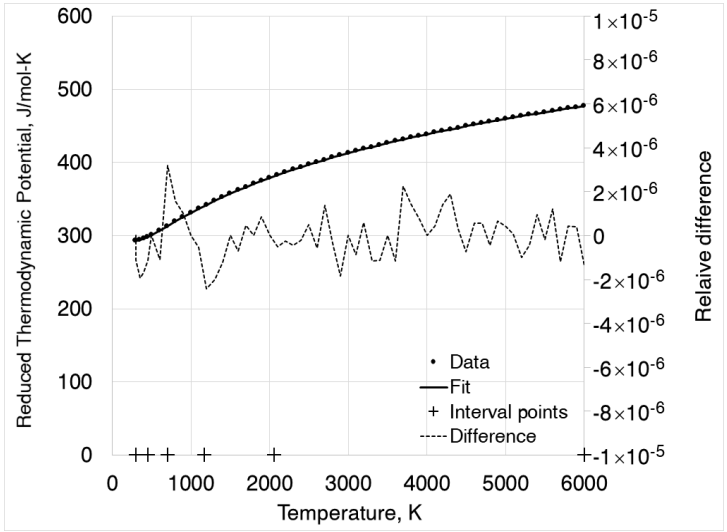
Dichlorofluoromethane ($CHCl_2F$) – C1H1Cl2F1(g)

MW = 102.923 g/mol – $\Delta_f H^o$(298.15 K) = -283.257 kJ/mol – S^o(298.15 K) = 293.26 J/mol-K – p^o = 0.1 MPa



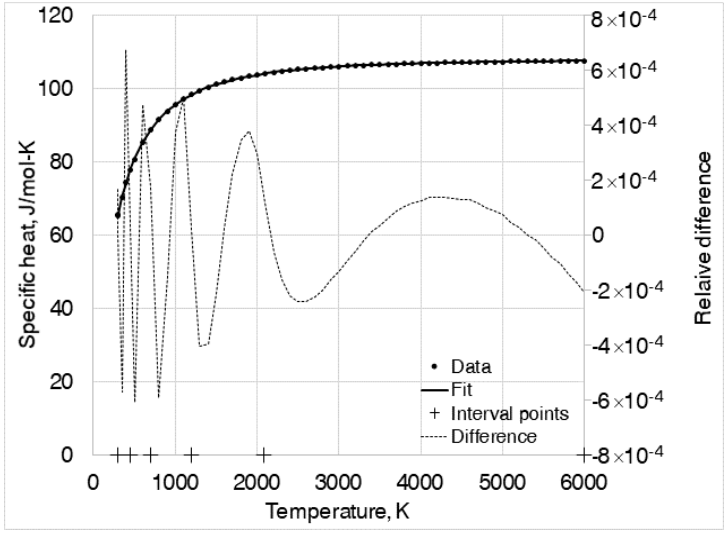


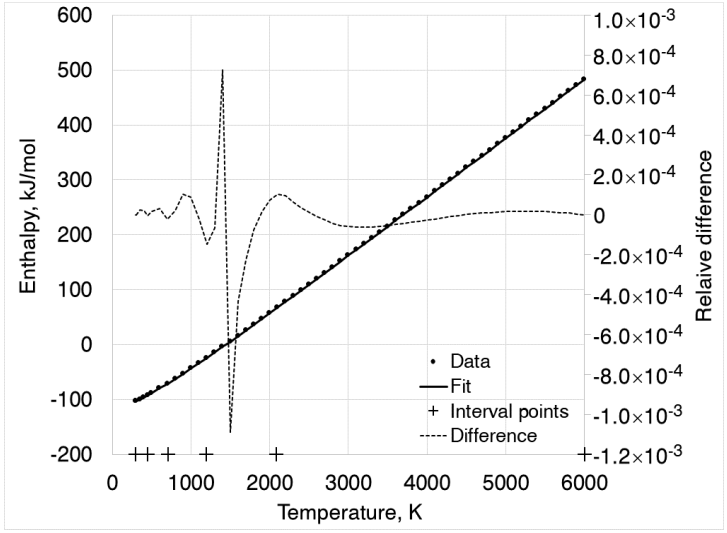


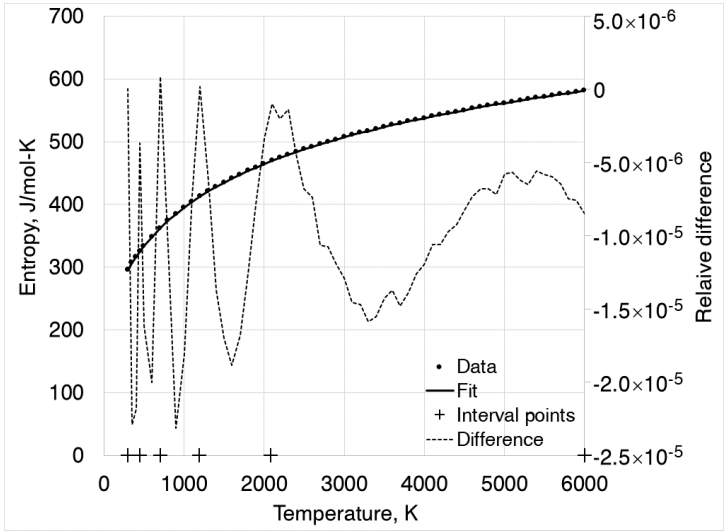


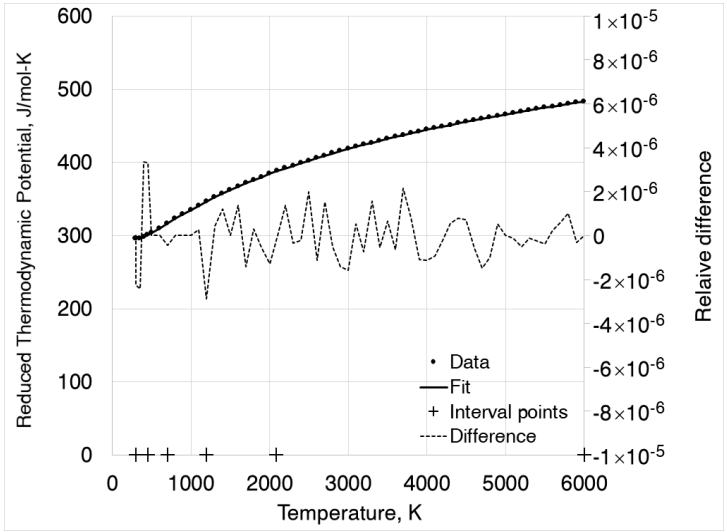
Trichloromethane ($CHCl_3$) – C1H1Cl3(g)

MW = 119.378 g/mol – $\Delta_f H^o$(298.15 K) = -103.177 kJ/mol – S^o(298.15 K) = 295.62 J/mol-K – p^o = 0.1 MPa



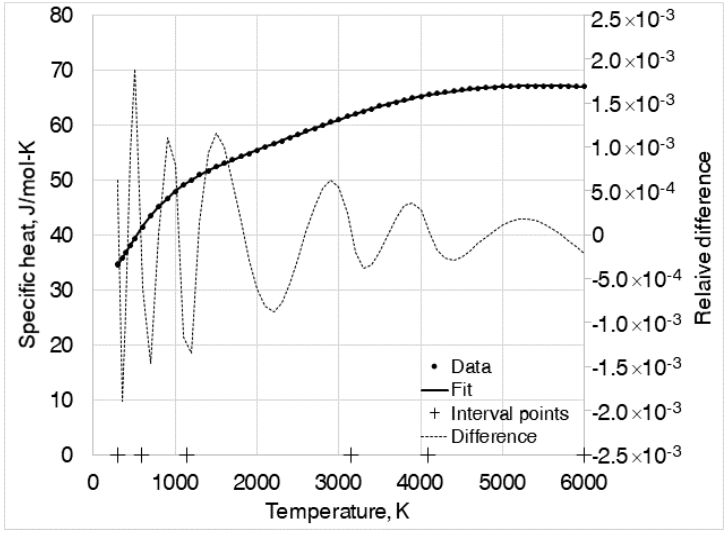


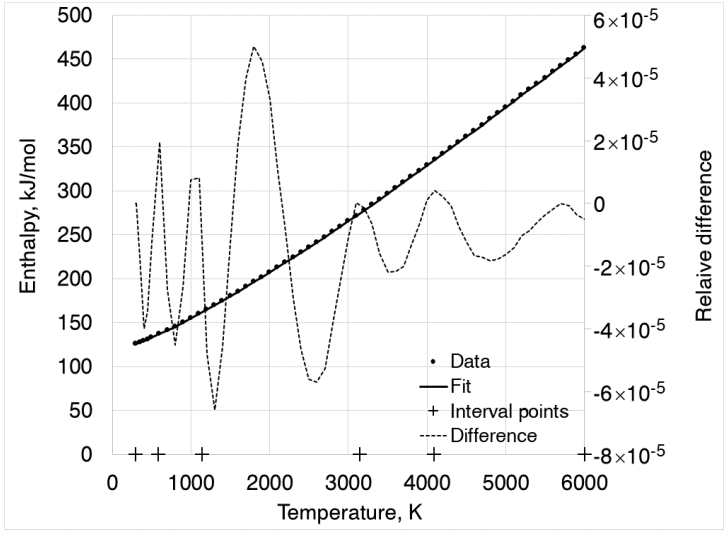


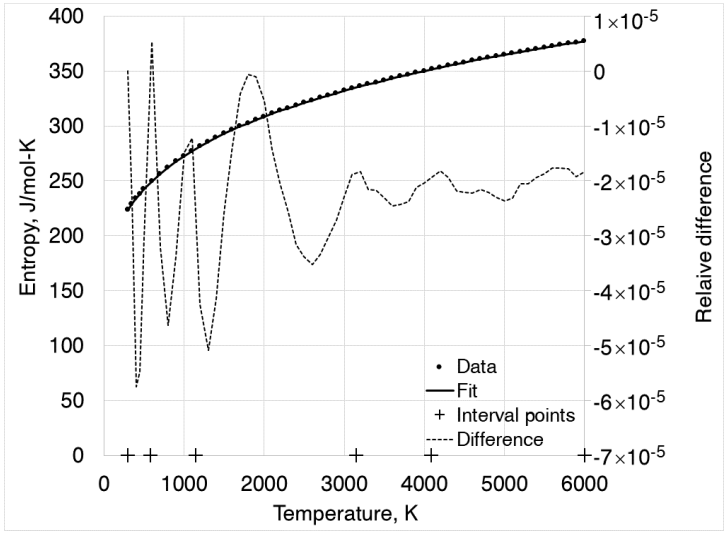


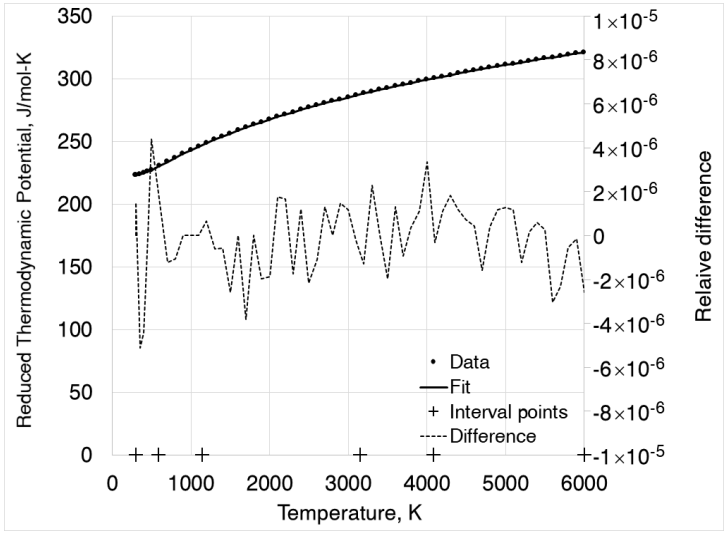
Fluoromethylene (CHF) – C1H1F1(g)

MW = 32.017 g/mol – $\Delta_f H^o$(298.15 K) = 125.52 kJ/mol – S^o(298.15 K) = 223.348 J/mol-K – p^o = 0.1 MPa



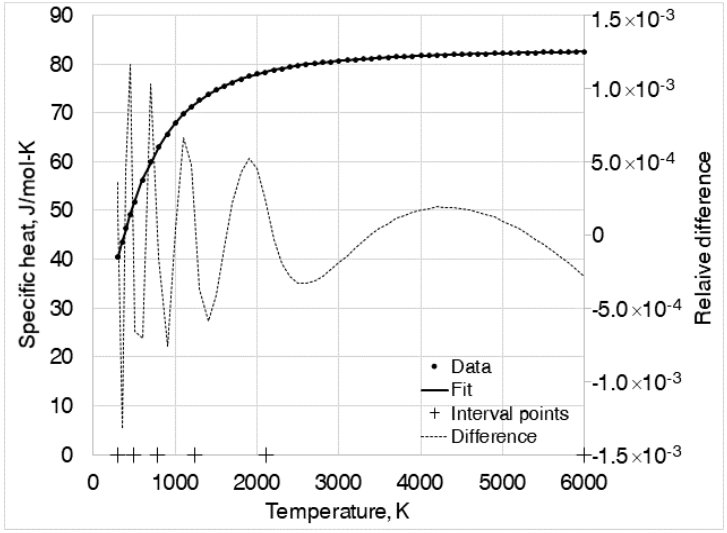


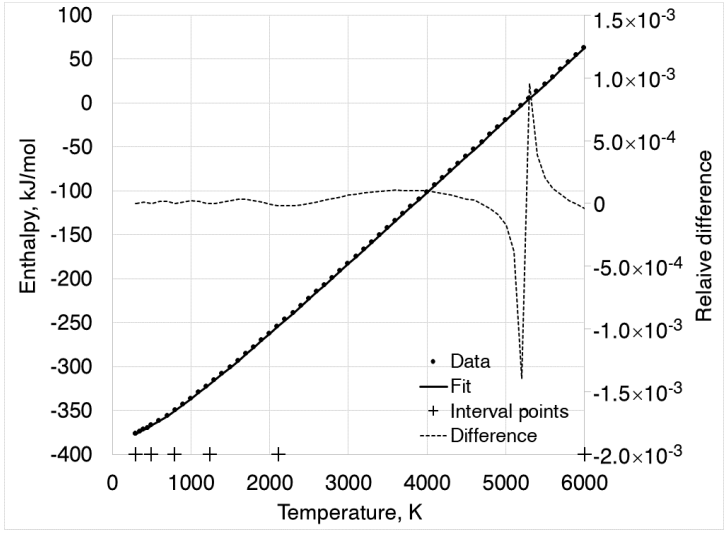


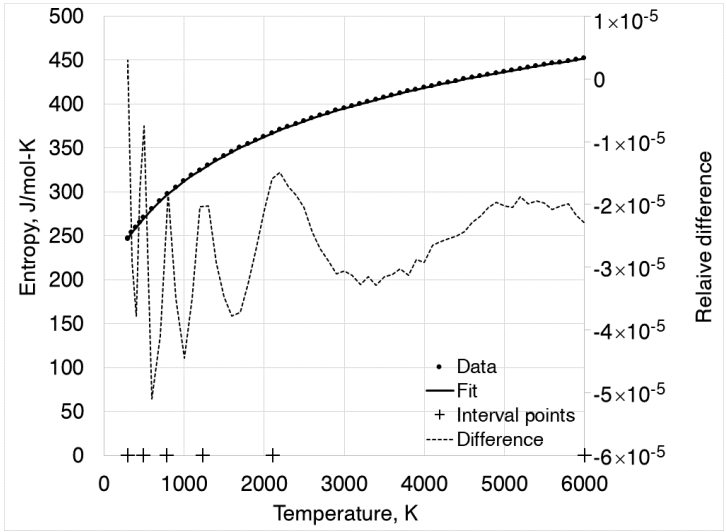


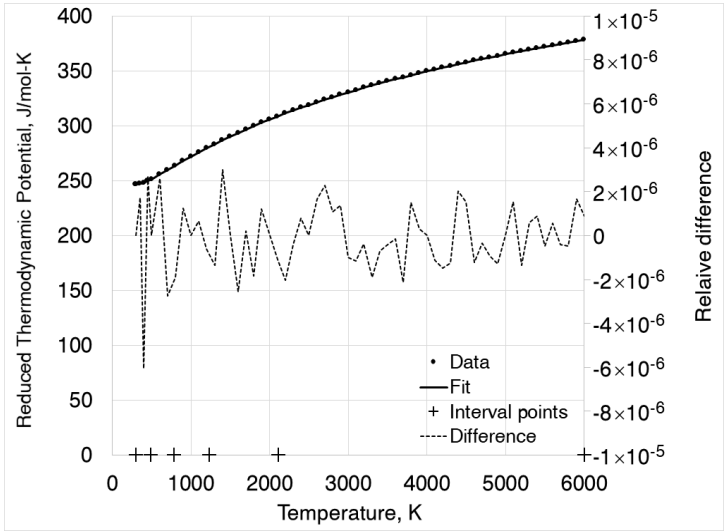
Formyl Fluoride (HCOF) – C1H1F1O1(g)

MW = 48.017 g/mol – $\Delta_f H^o$(298.15 K) = -376.56 kJ/mol – S^o(298.15 K) = 246.52 J/mol-K – p^o = 0.1 MPa



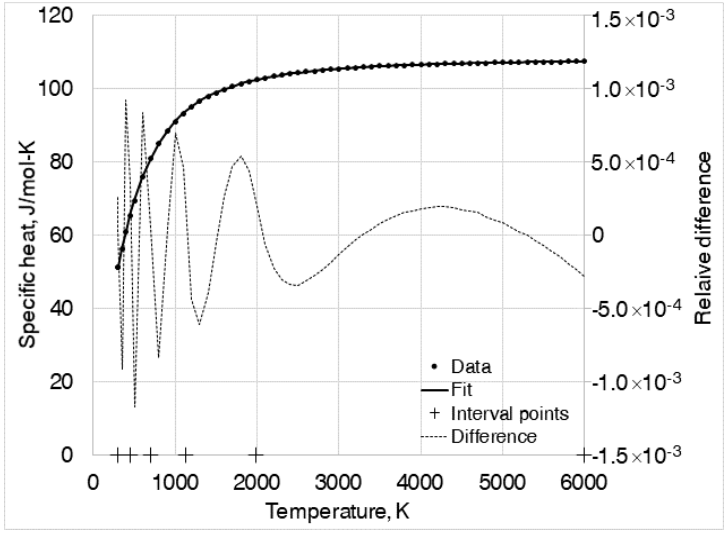


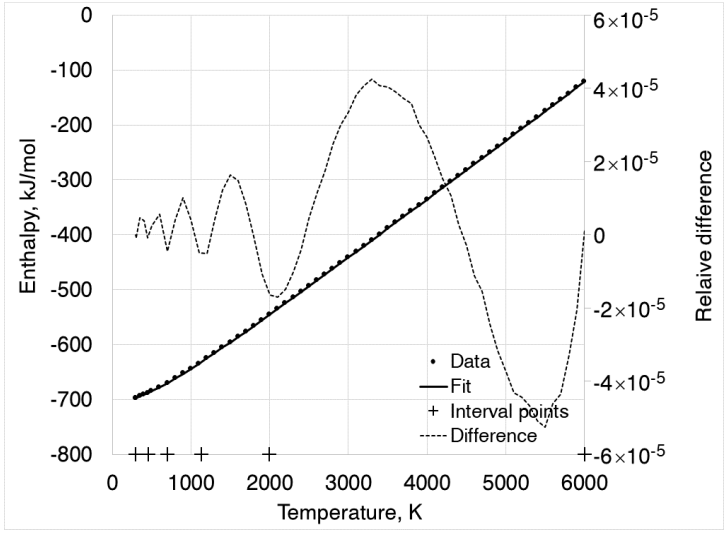


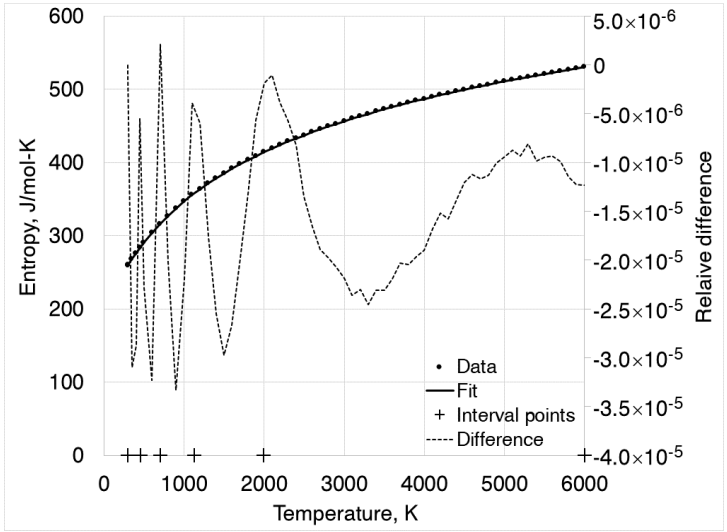


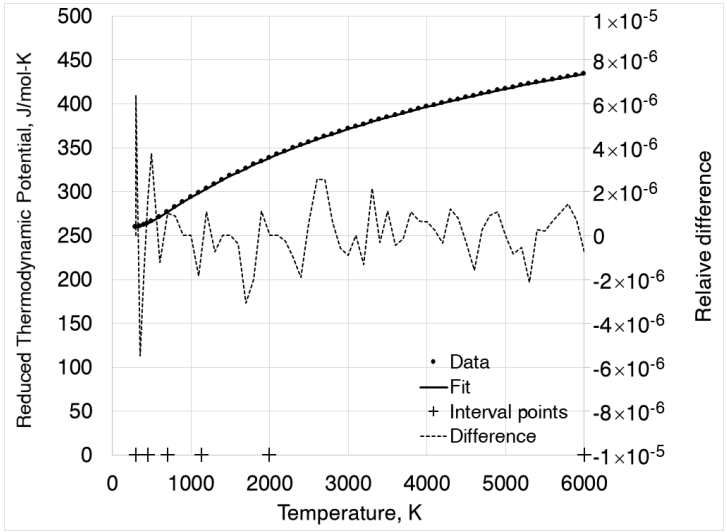
Trifluoromethane (CHF_3) – C1H1F3(g)

MW = 70.014 g/mol – $\Delta_f H^o$(298.15 K) = -697.054 kJ/mol – S^o(298.15 K) = 259.657 J/mol-K – p^o = 0.1 MPa



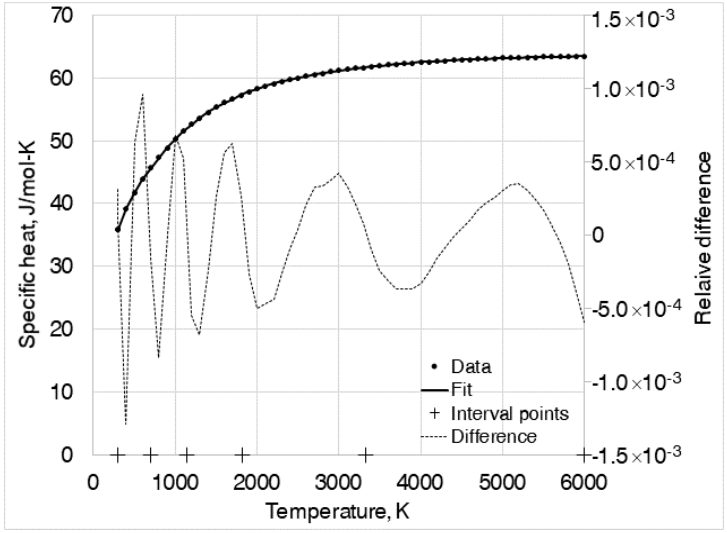


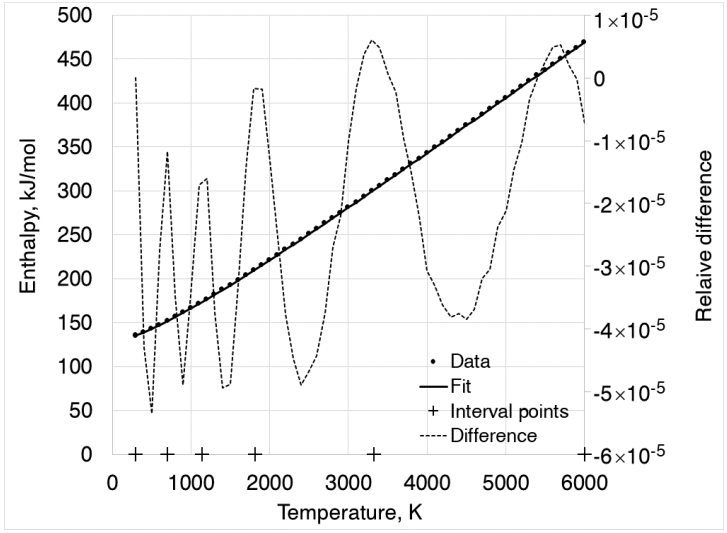


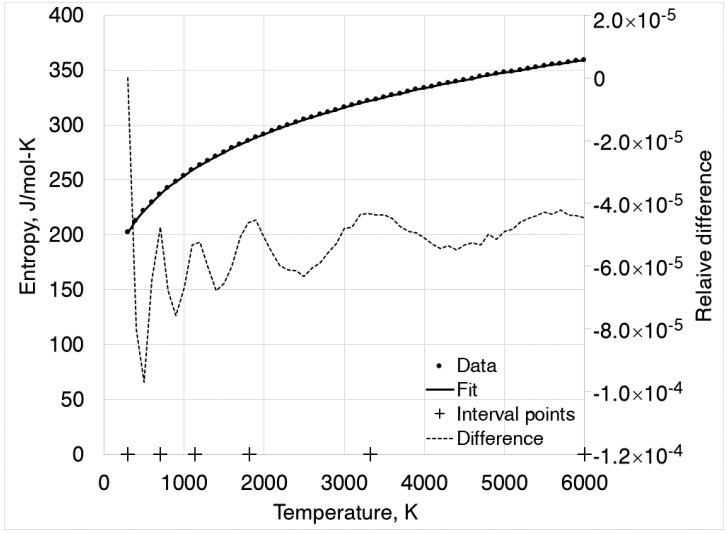


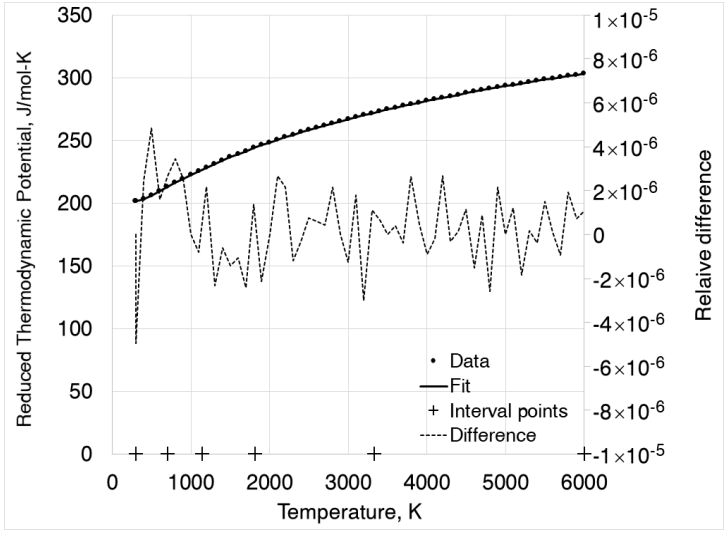
Hydrogen Cyanide (HCN) – C1H1N1(g)

MW = 27.026 g/mol – $\Delta_f H^o$(298.15 K) = 135.143 kJ/mol – S^o(298.15 K) = 201.828 J/mol-K – p^o = 0.1 MPa



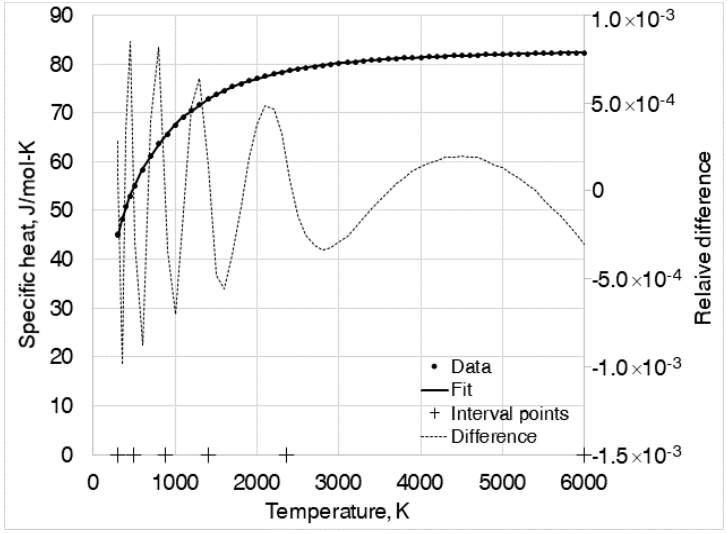


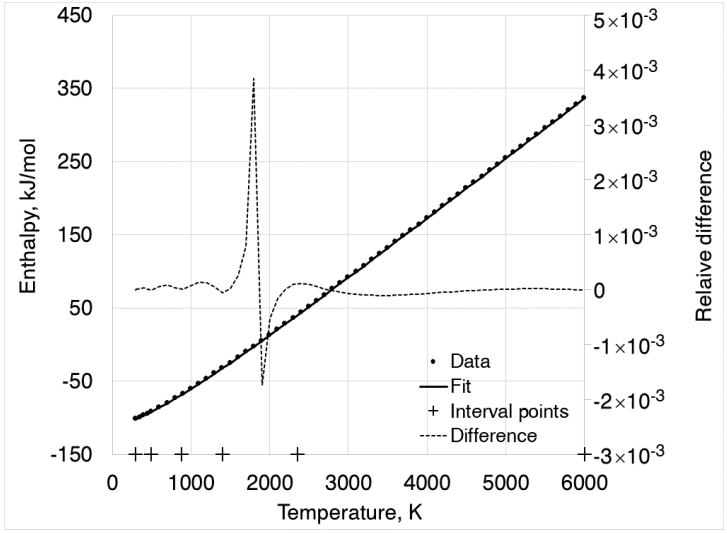


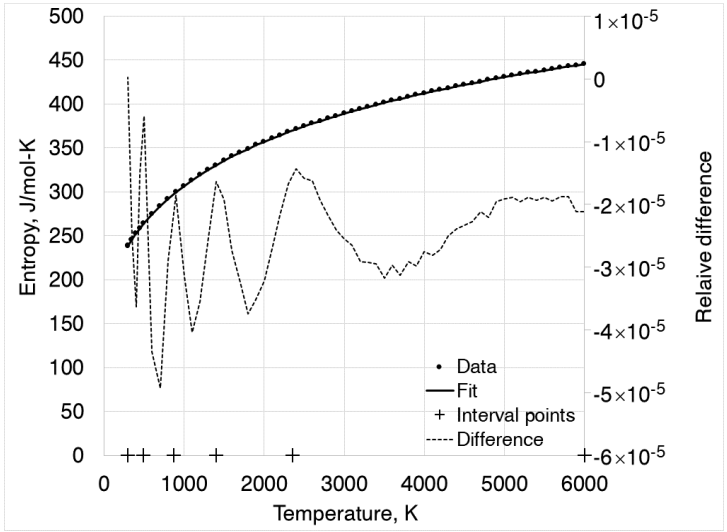


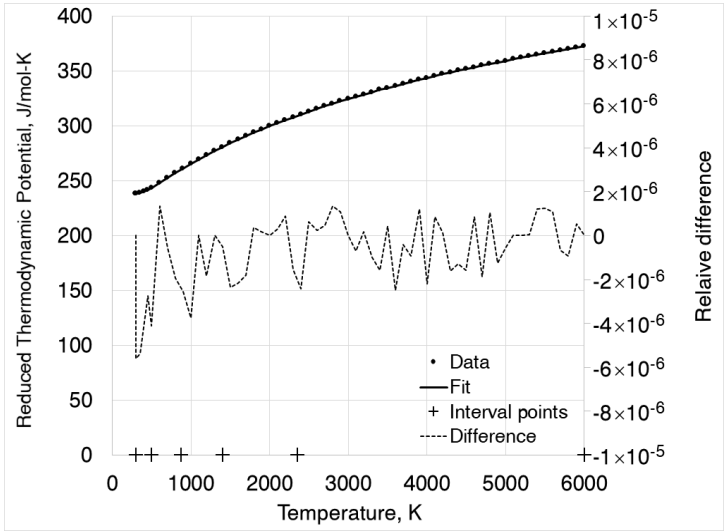
Hydrogen Isocyanate (HNCO) – C1H1N1O1(g)

MW = 43.025 g/mol – $\Delta_f H^o$(298.15 K) = -101.671 kJ/mol – S^o(298.15 K) = 238.228 J/mol-K – p^o = 0.1 MPa



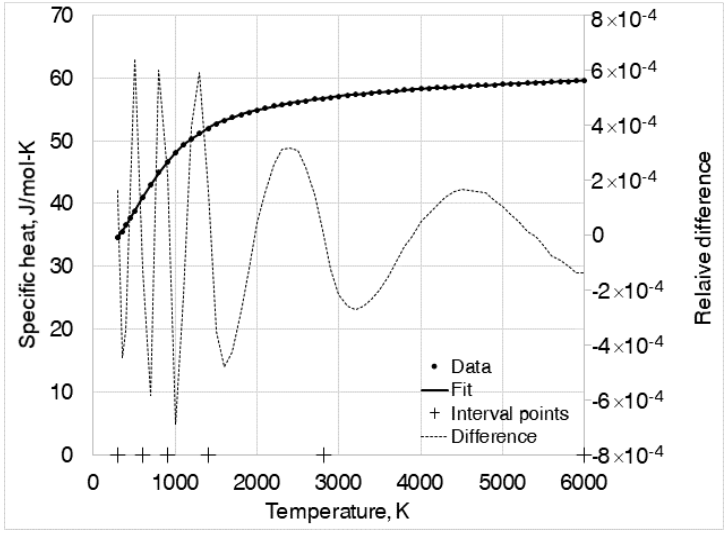


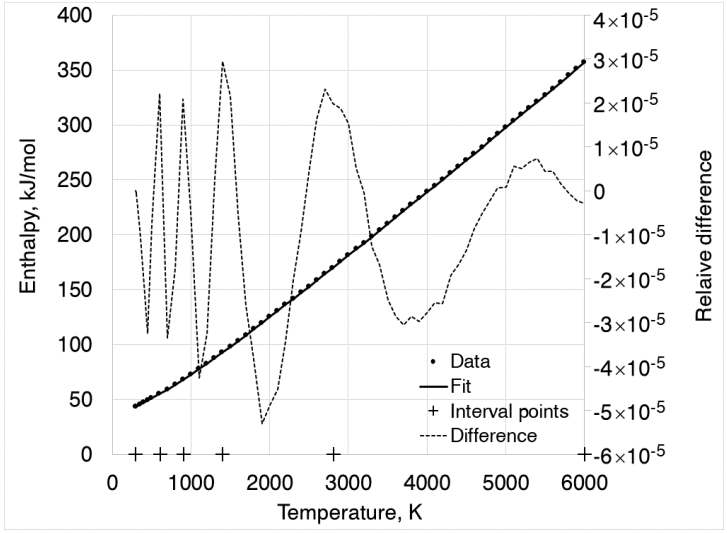


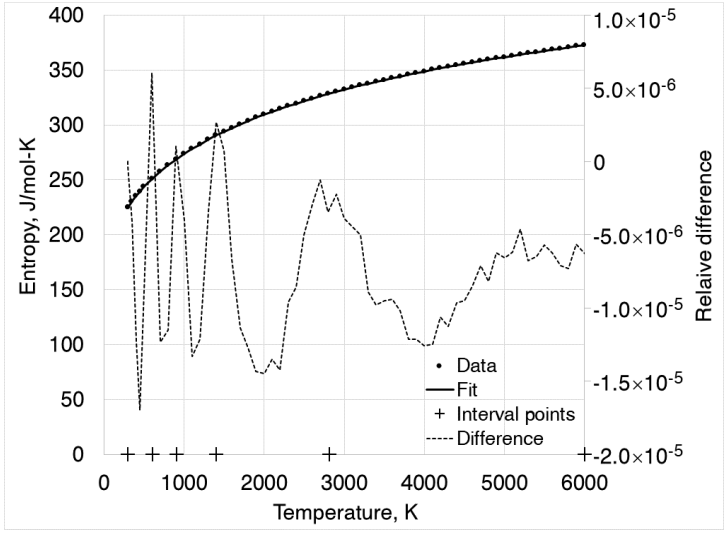


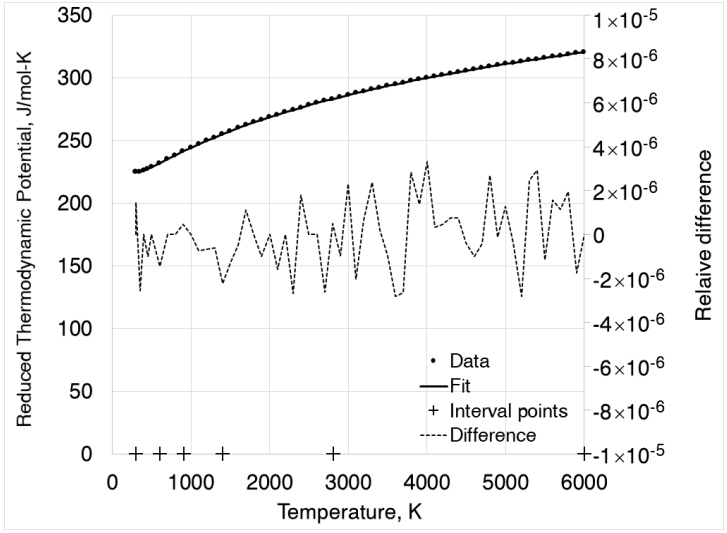
Formyl (HCO) – C1H1O1(g)

MW = 29.018 g/mol – $\Delta_f H^o$(298.15 K) = 43.514 kJ/mol – S^o(298.15 K) = 224.649 J/mol-K – p^o = 0.1 MPa



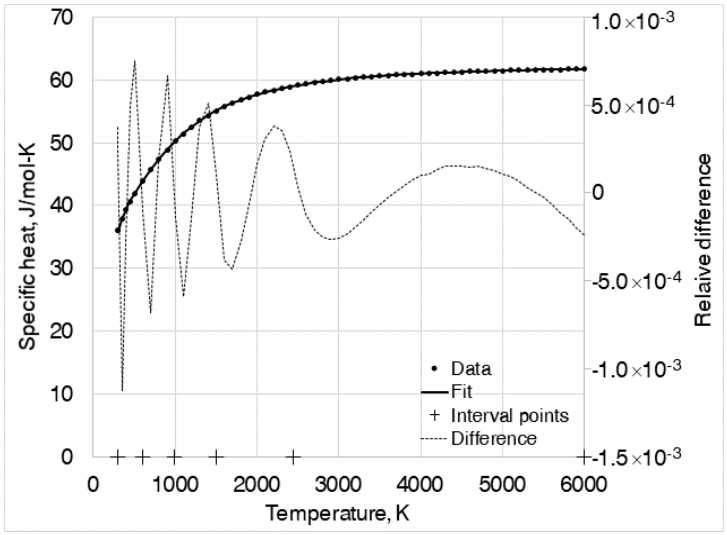


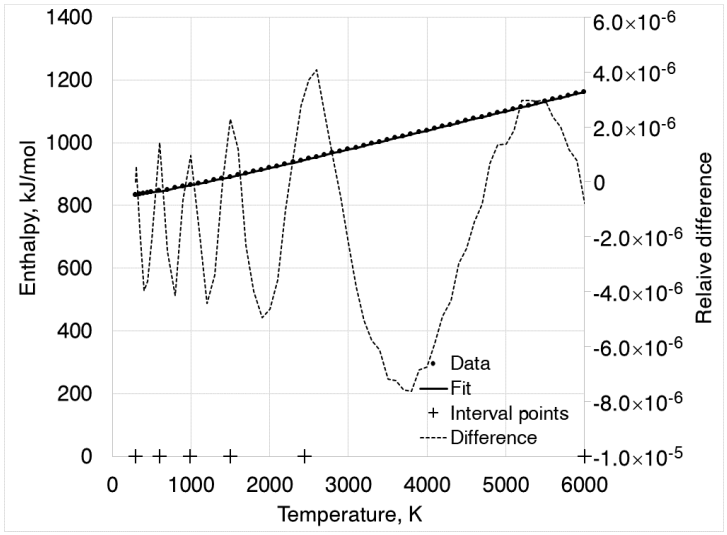


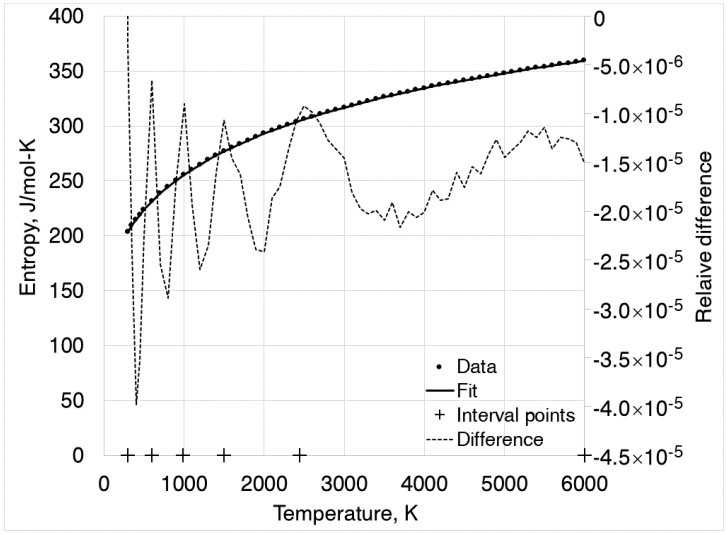


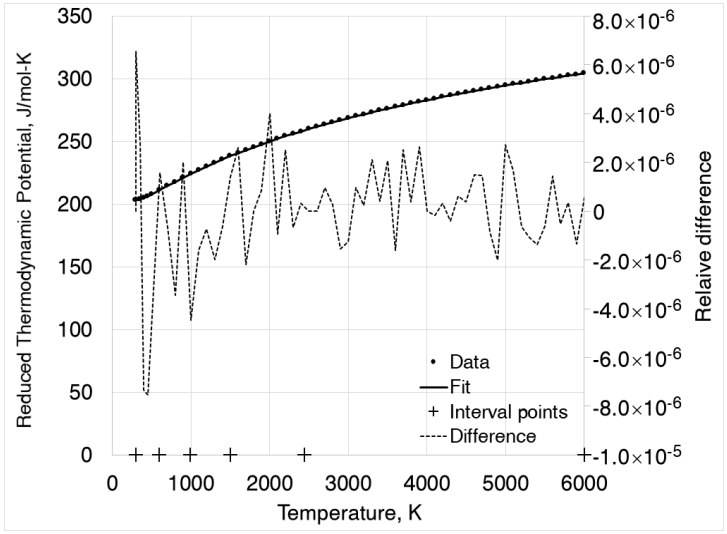
Formyl, Ion (HCO^+) – C1H1O1+(g)

MW = 29.018 g/mol – $\Delta_f H^o$(298.15 K) = 833.034 kJ/mol – S^o(298.15 K) = 203.417 J/mol-K – p^o = 0.1 MPa



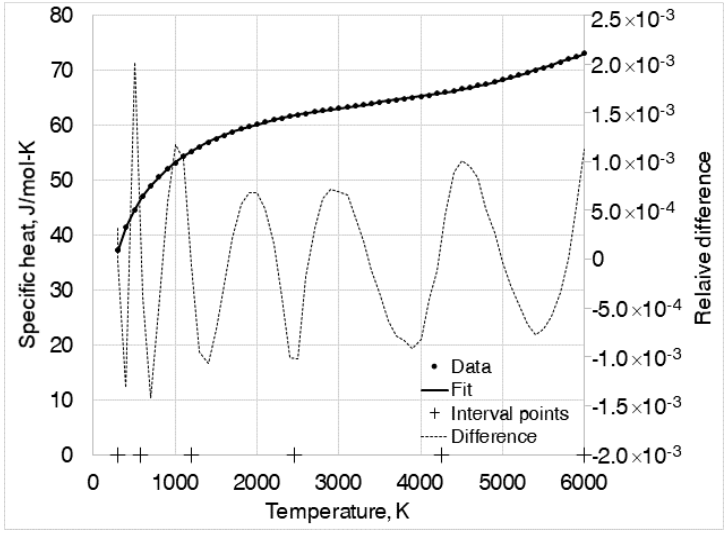


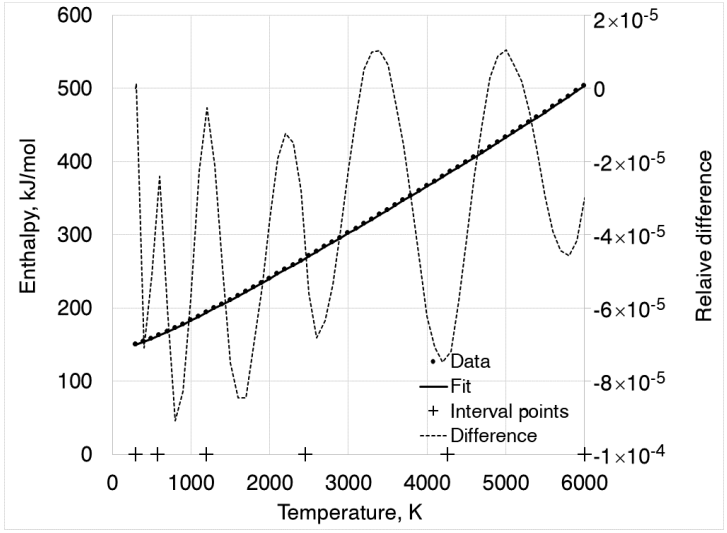


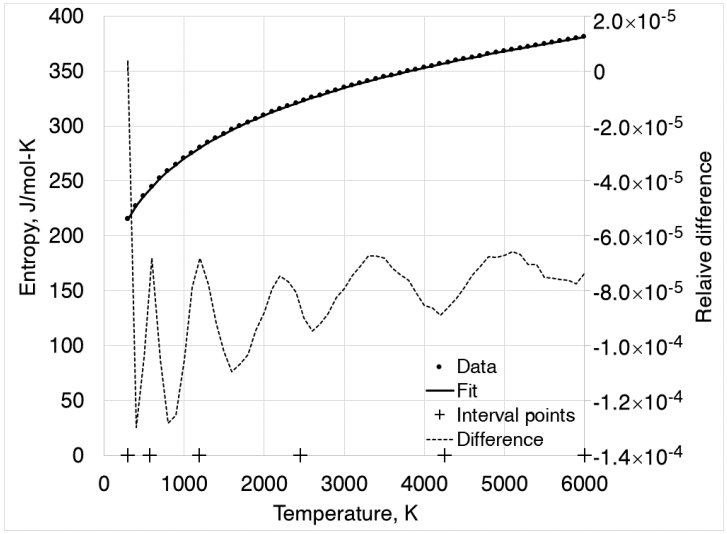


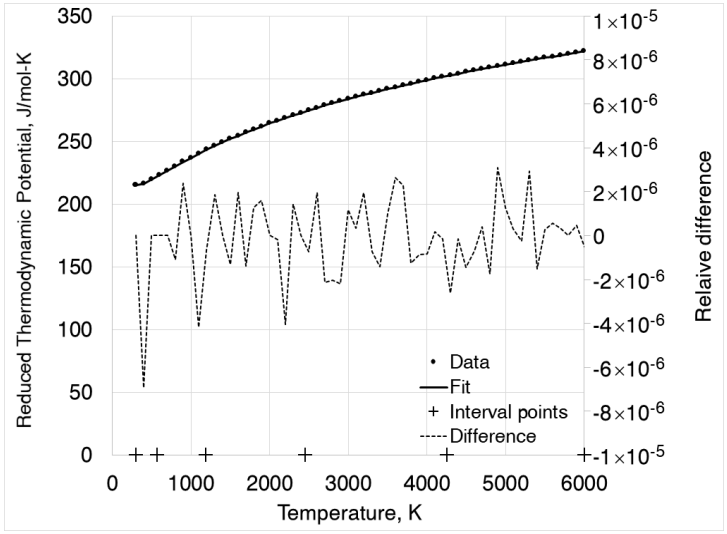
Methinophosphide (CHP) – C1H1P1(g)

MW = 43.993 g/mol – $\Delta_f H^o$(298.15 K) = 149.9 kJ/mol – S^o(298.15 K) = 215.062 J/mol-K – p^o = 0.1 MPa



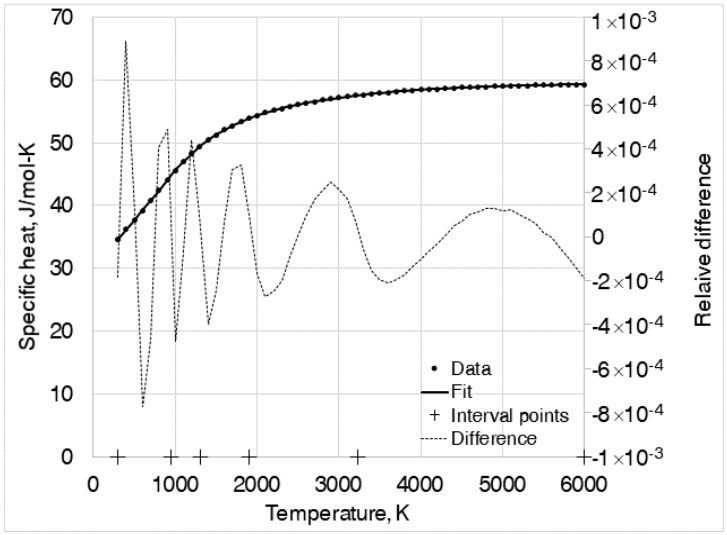


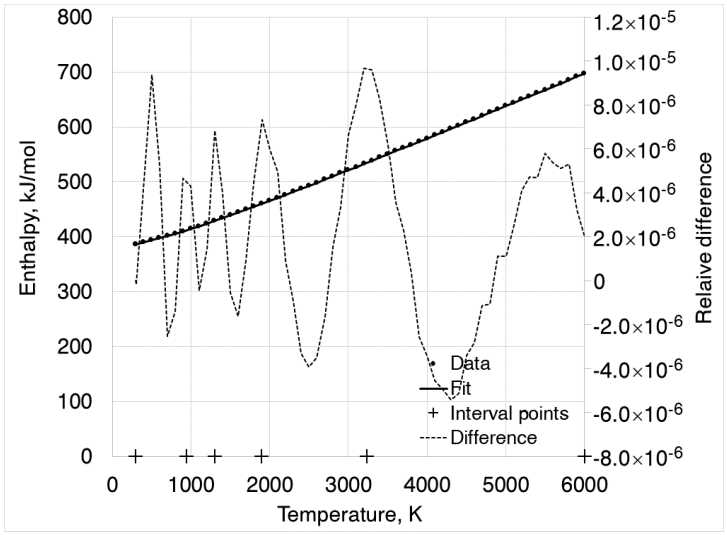


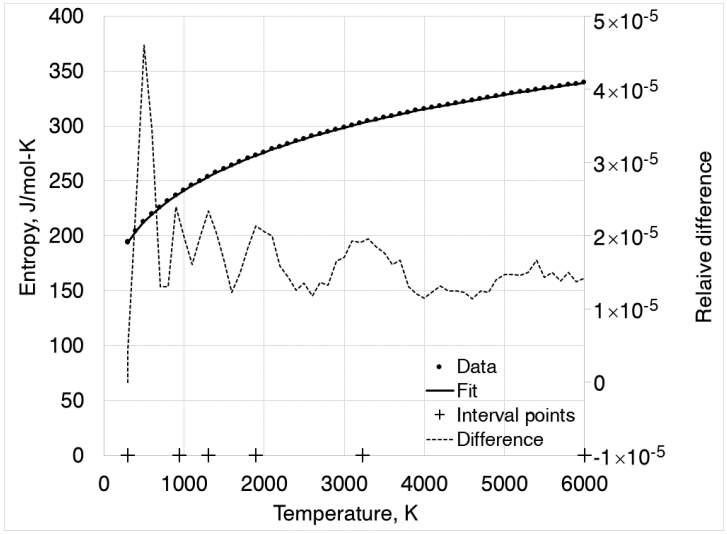


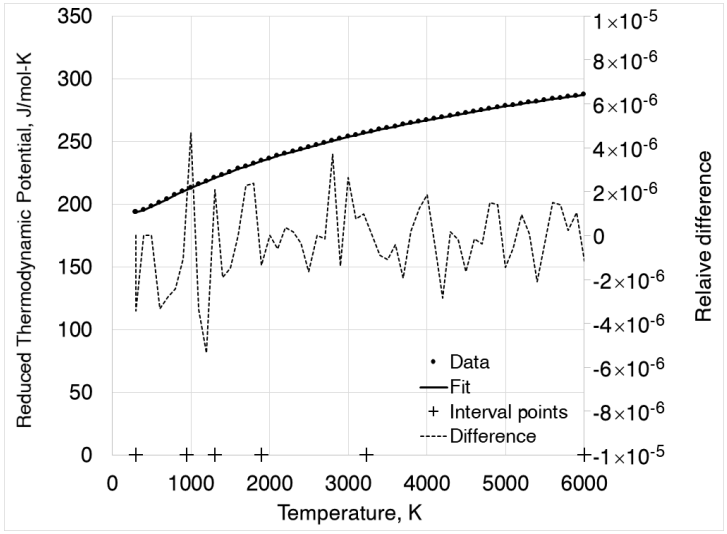
Methylene (CH_2) – C1H2(g)

MW = 14.027 g/mol – $\Delta_f H^o$(298.15 K) = 386.392 kJ/mol – S^o(298.15 K) = 193.931 J/mol-K – p^o = 0.1 MPa



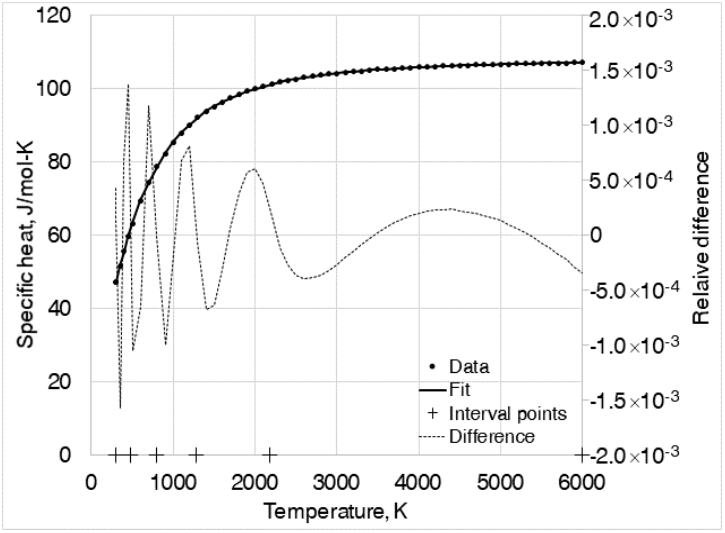


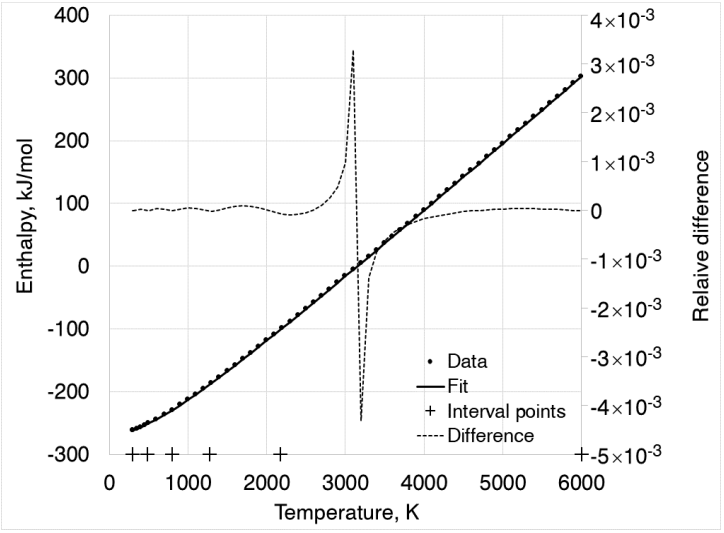


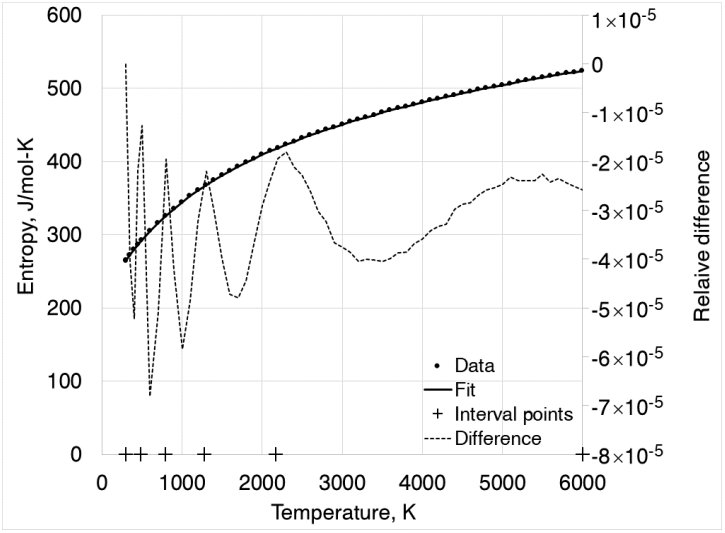


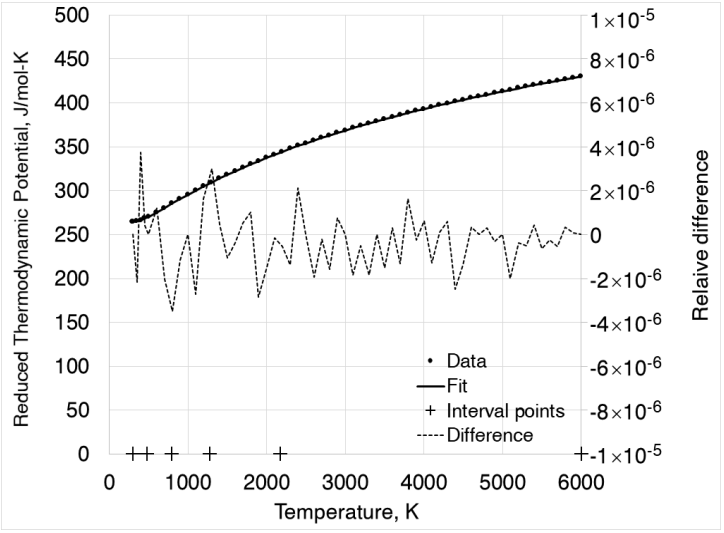
Chlorofluoromethane (CH_2ClF) – C1H2Cl1F1(g)

MW = 68.478 g/mol – $\Delta_f H^o$(298.15 K) = -261.918 kJ/mol – S^o(298.15 K) = 264.421 J/mol-K – p^o = 0.1 MPa



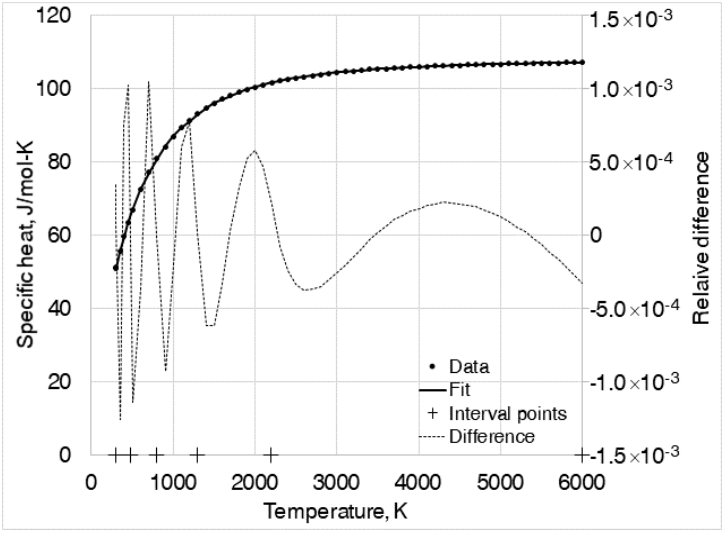


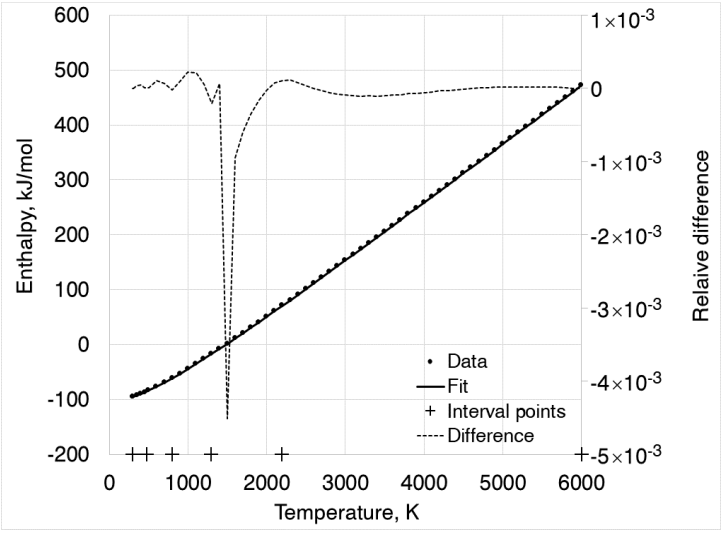


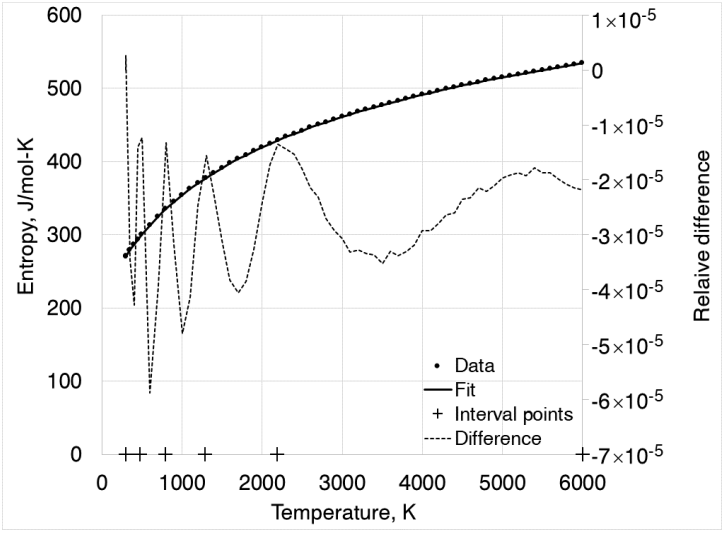


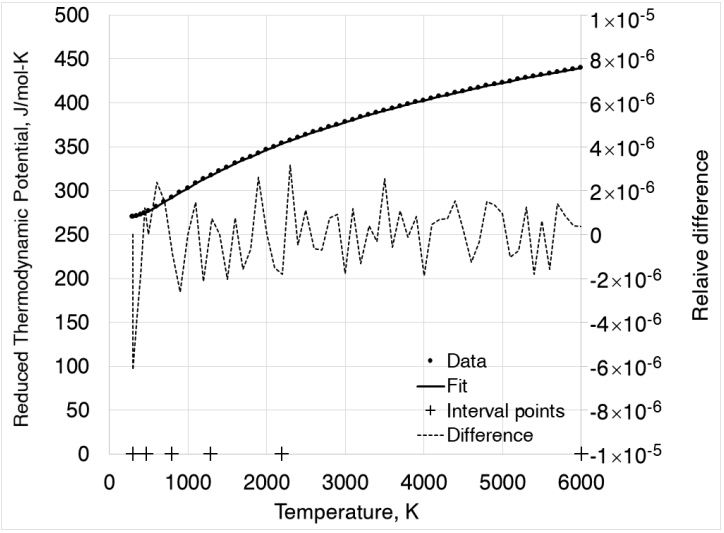
Dichloromethane (CH_2Cl_2) – C1H2Cl2(g)

MW = 84.933 g/mol – $\Delta_f H^o$(298.15 K) = -95.521 kJ/mol – S^o(298.15 K) = 270.293 J/mol-K – p^o = 0.1 MPa



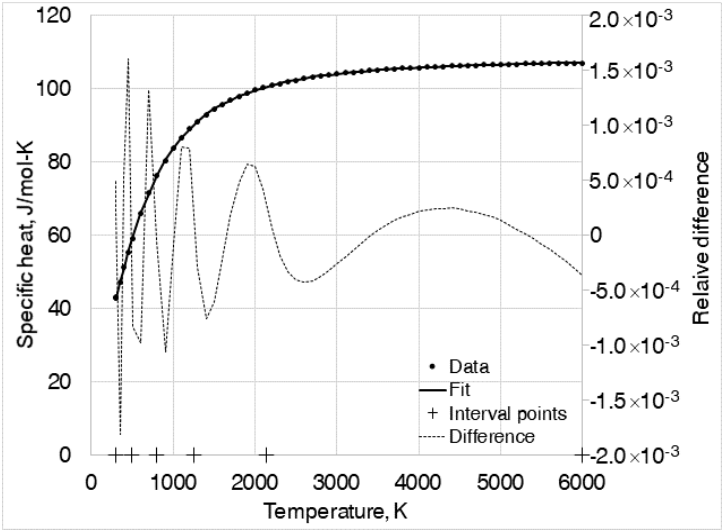


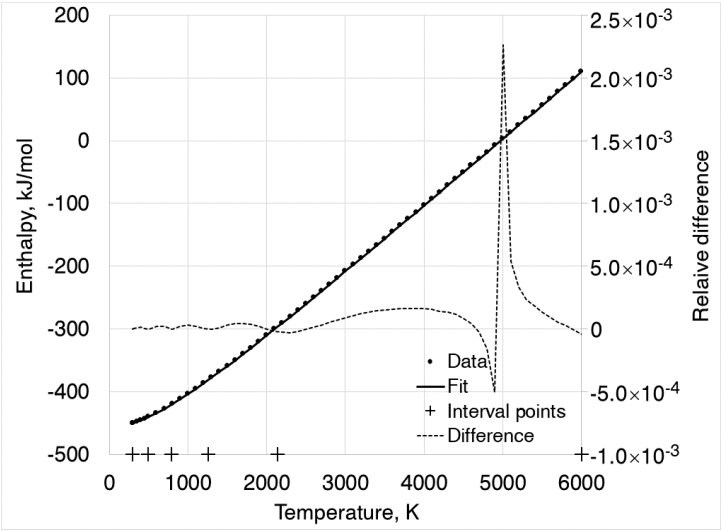


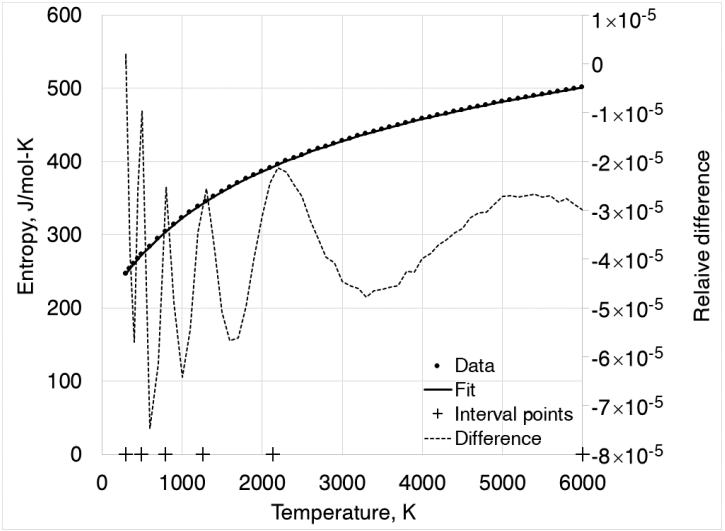


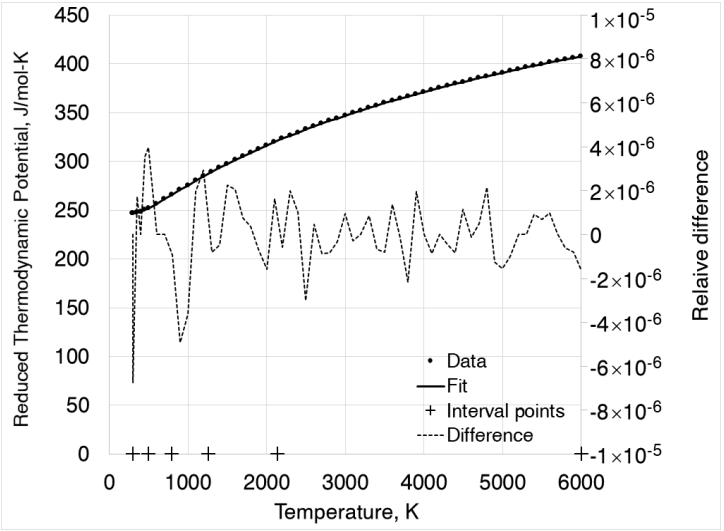
Difluoromethane (CH_2F_2) – C1H2F2(g)

MW = 52.024 g/mol – $\Delta_f H^o$(298.15 K) = -450.659 kJ/mol – S^o(298.15 K) = 246.698 J/mol-K – p^o = 0.1 MPa



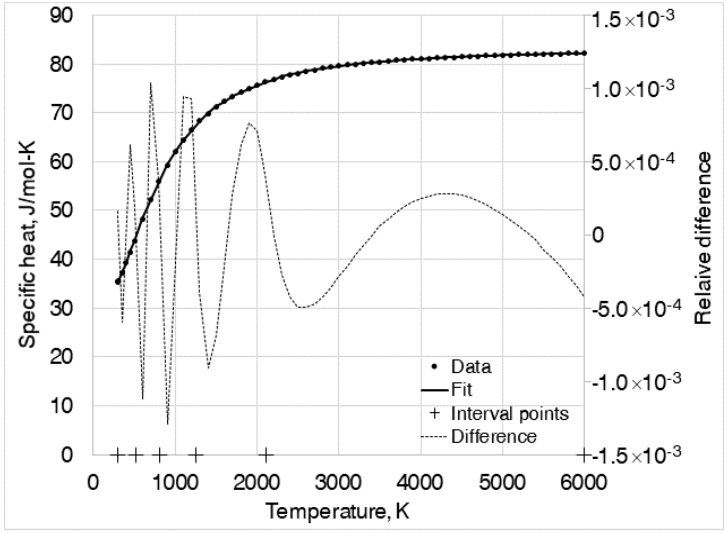


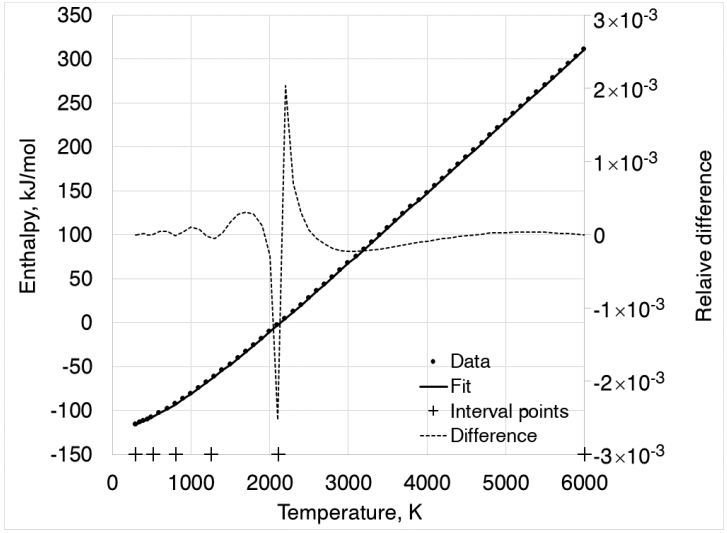


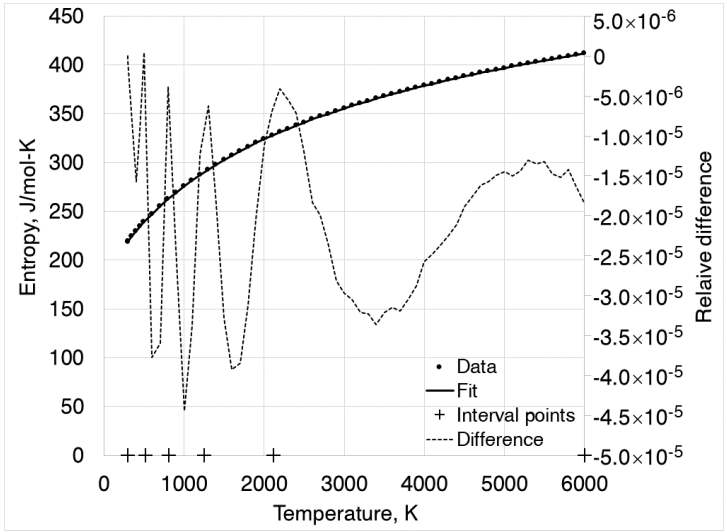


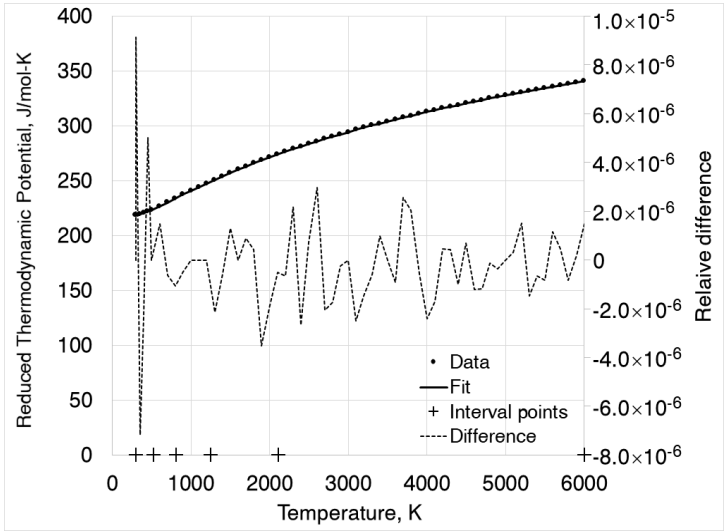
Formaldehyde (H_2CO) – C1H2O1(g)

MW = 30.026 g/mol – $\Delta_f H^o$(298.15 K) = -115.897 kJ/mol – S^o(298.15 K) = 218.95 J/mol-K – p^o = 0.1 MPa



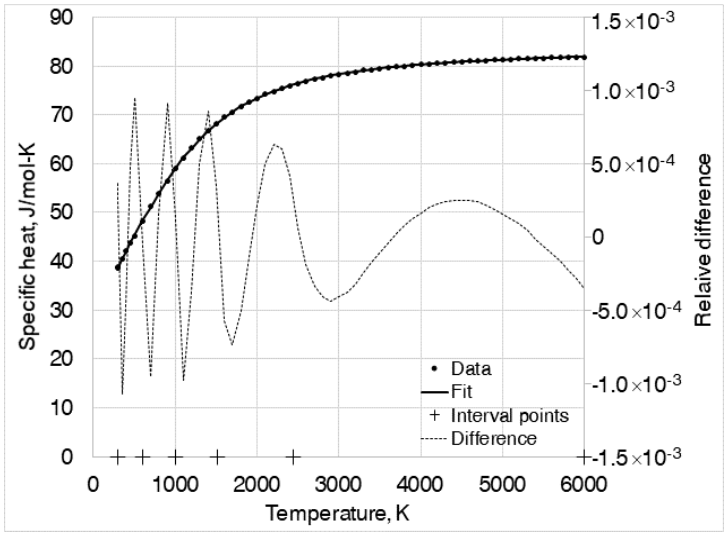


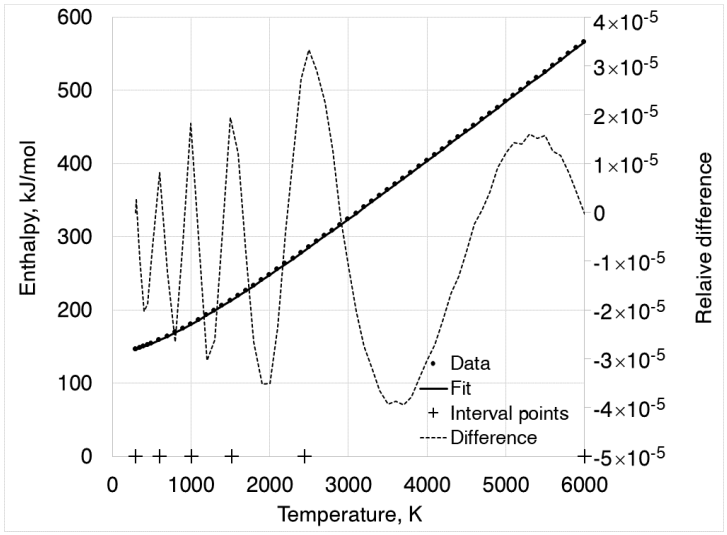


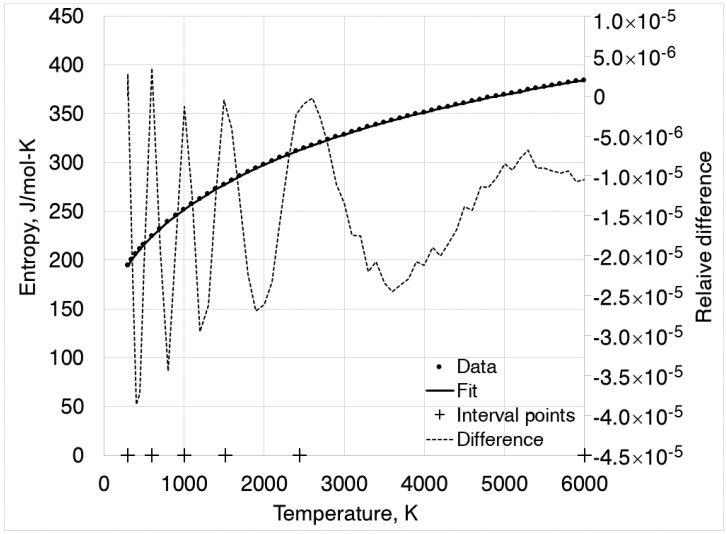


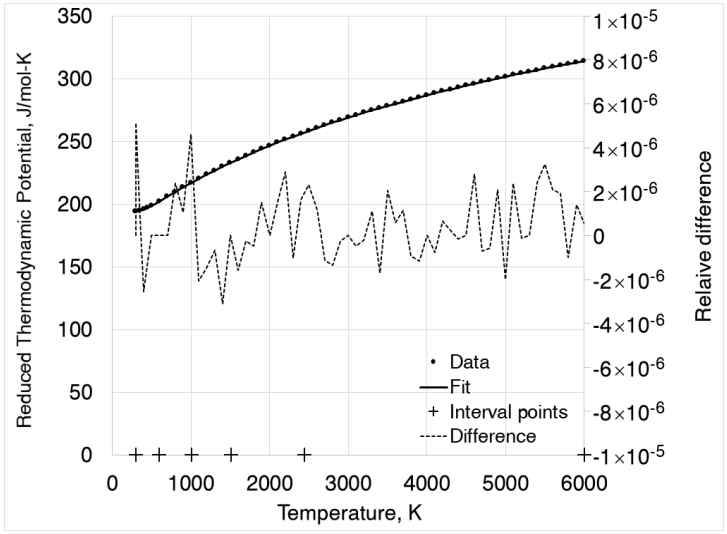
Methyl (CH_3) – C1H3(g)

MW = 15.035 g/mol – $\Delta_f H^o$(298.15 K) = 145.687 kJ/mol – S^o(298.15 K) = 194.17 J/mol-K – p^o = 0.1 MPa



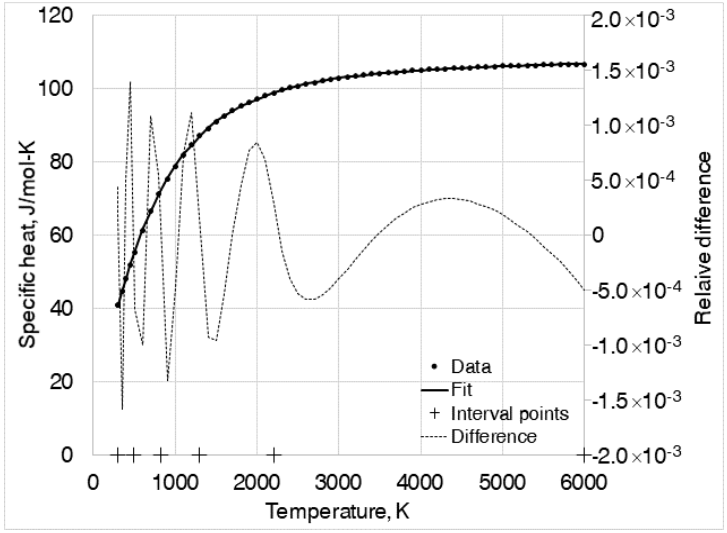


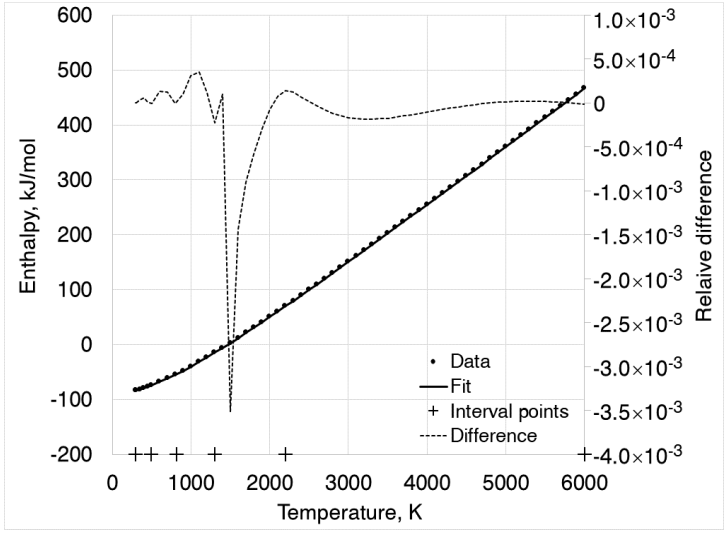


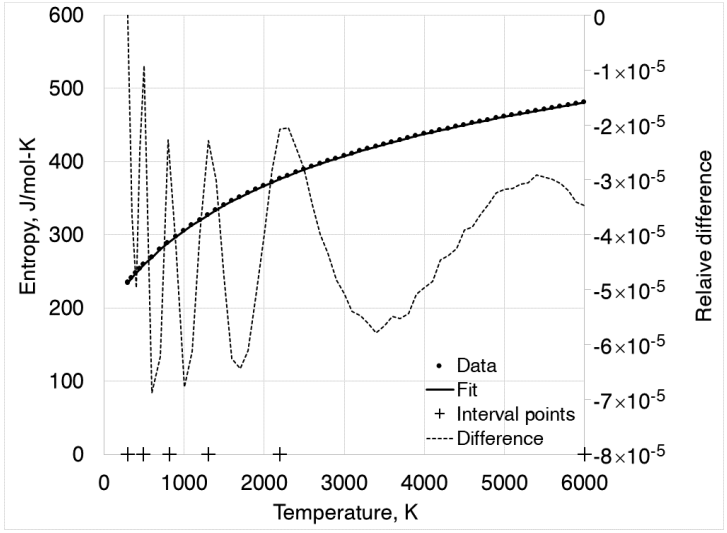


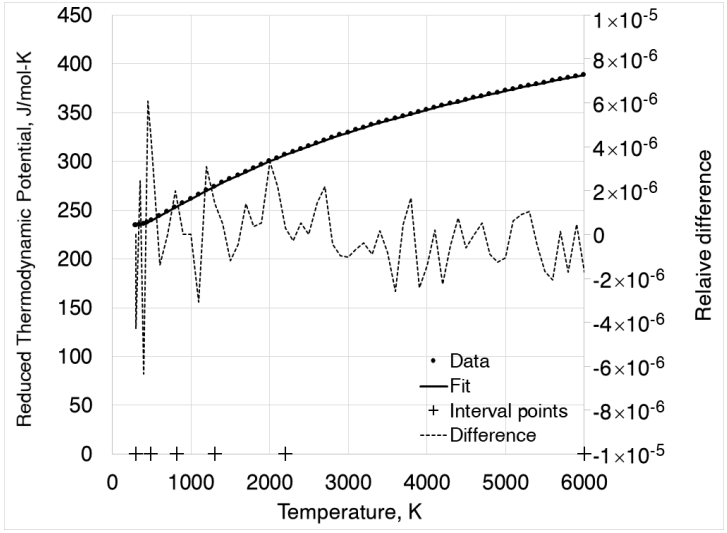
Chloromethane (CH_3Cl) – C1H3Cl1(g)

MW = 50.488 g/mol – $\Delta_f H^o$(298.15 K) = -83.68 kJ/mol – S^o(298.15 K) = 234.367 J/mol-K – p^o = 0.1 MPa



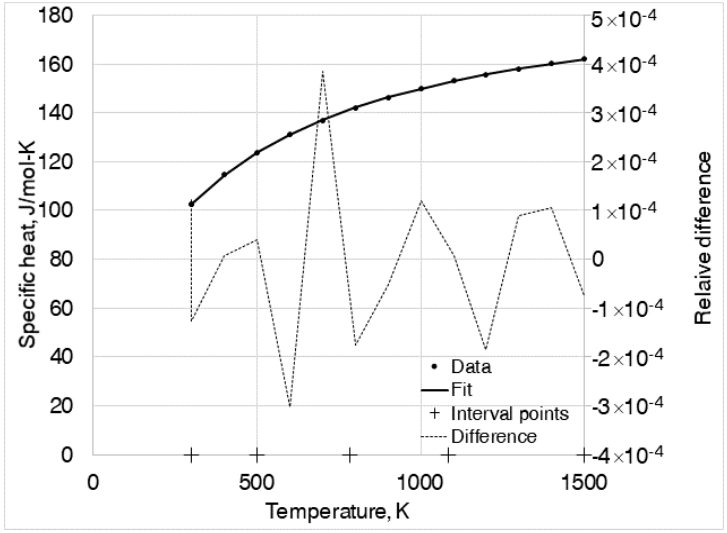


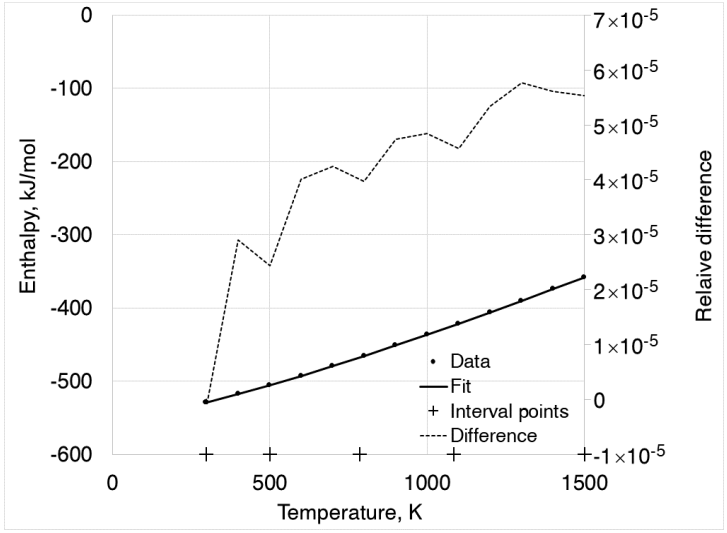


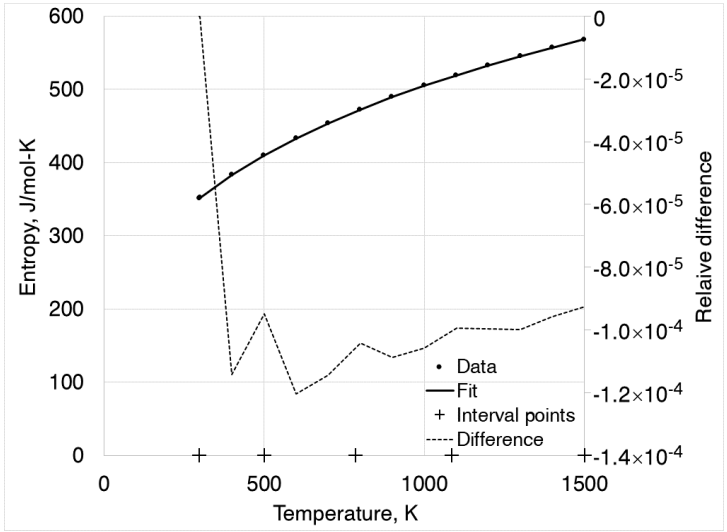


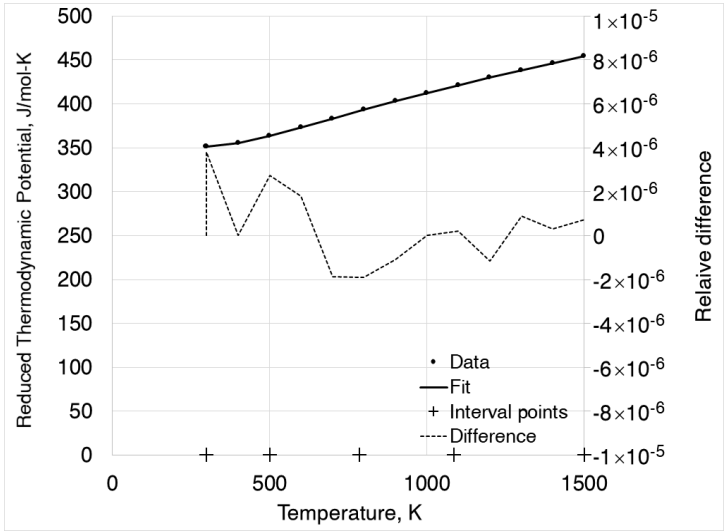
Trichloromethylsilane ($SiCH_3Cl_3$) – C1H3Cl3Si1(g)

MW = 149.479 g/mol – $\Delta_f H^o$(298.15 K) = -528.858 kJ/mol – S^o(298.15 K) = 351.147 J/mol-K – p^o = 0.1 MPa



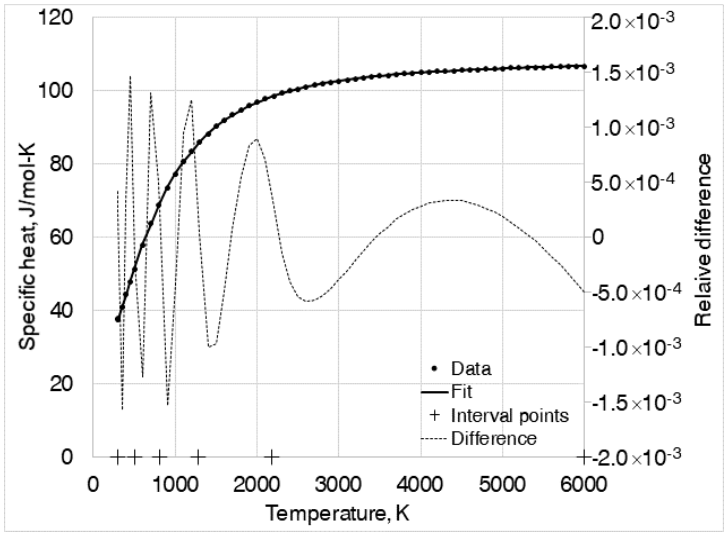


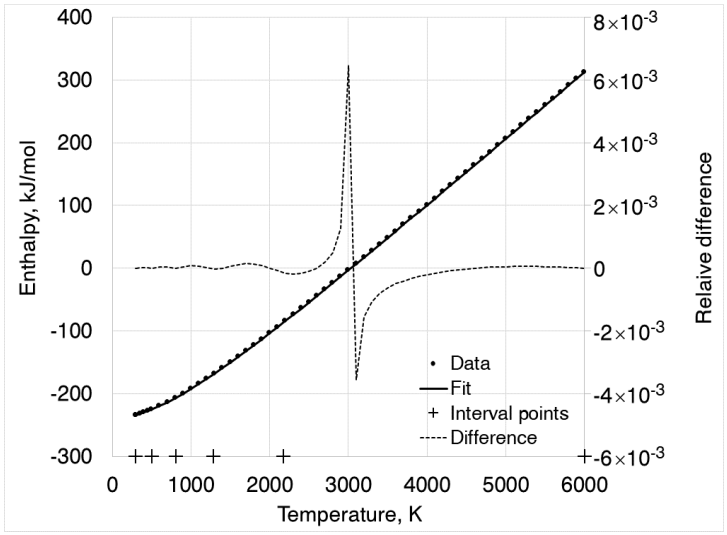


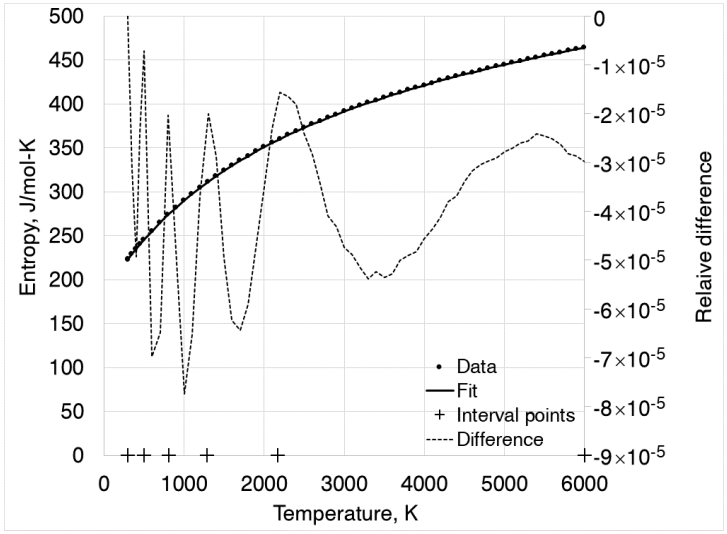


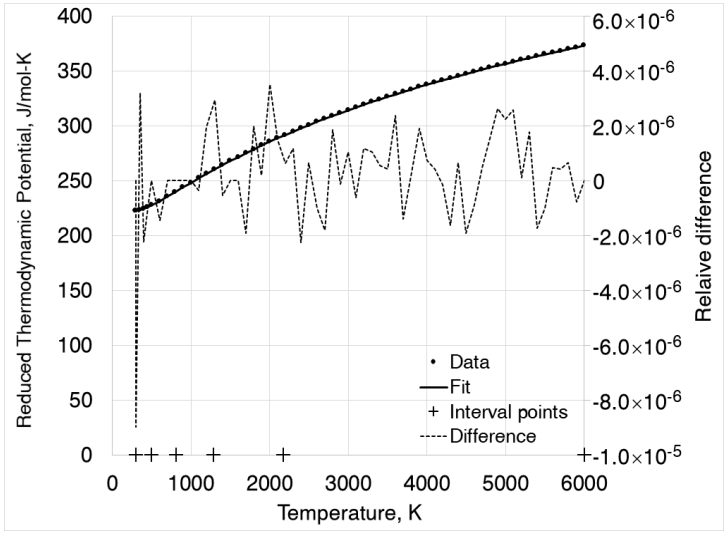
Fluoromethane (CH_3F) – C1H3F1(g)

MW = 34.033 g/mol – $\Delta_f H^o$(298.15 K) = -234.304 kJ/mol – S^o(298.15 K) = 222.843 J/mol-K – p^o = 0.1 MPa



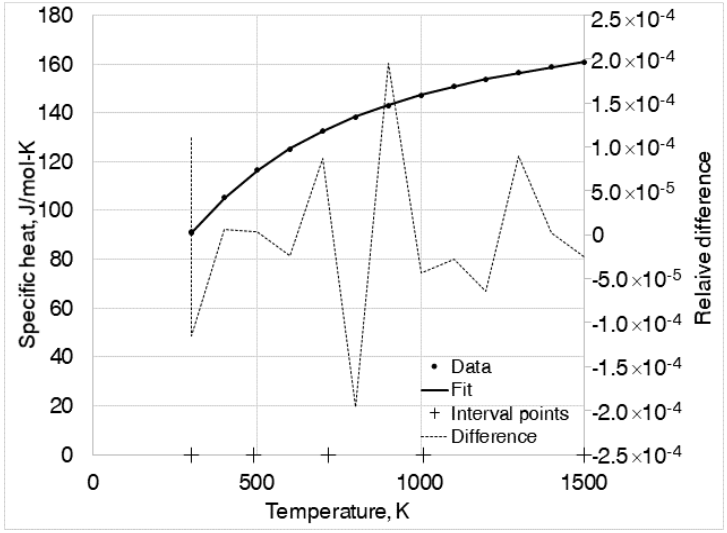


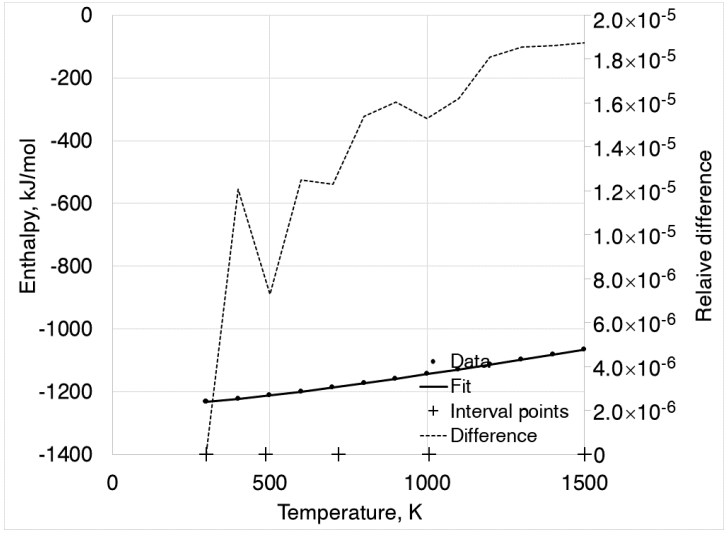


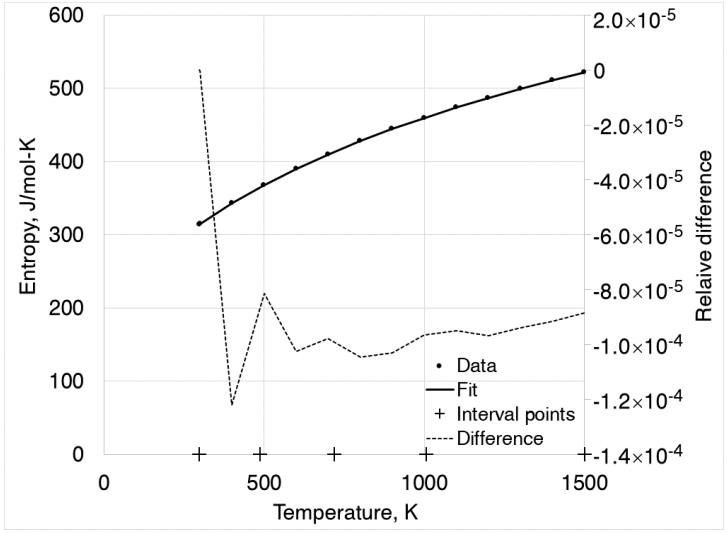


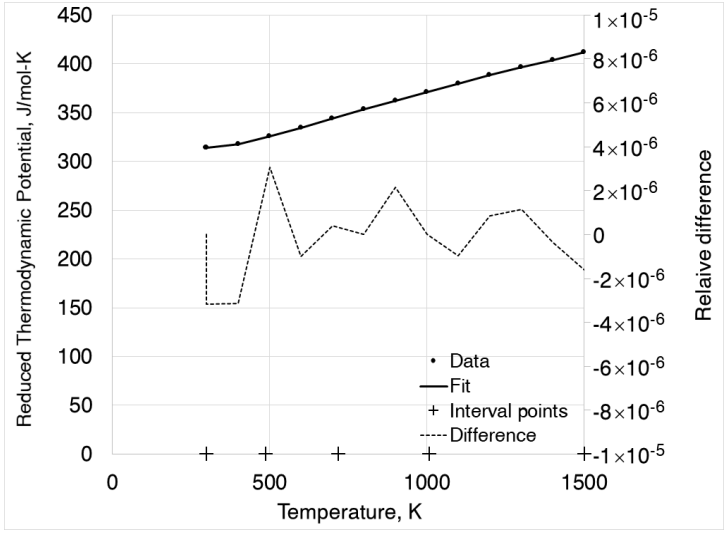
Trifluoromethylsilane ($SiCH_3F_3$) – C1H3F3Si1(g)

MW = 100.116 g/mol – $\Delta_f H^o$(298.15 K) = -1232.711 kJ/mol – S^o(298.15 K) = 314.202 J/mol-K – p^o = 0.1 MPa



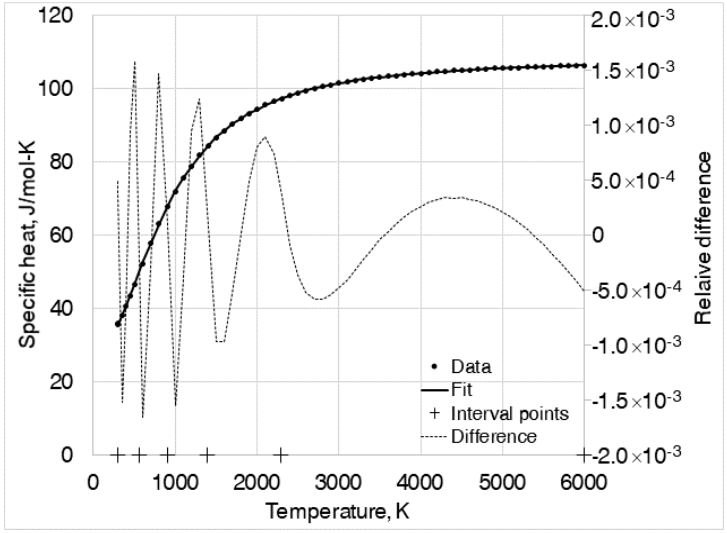


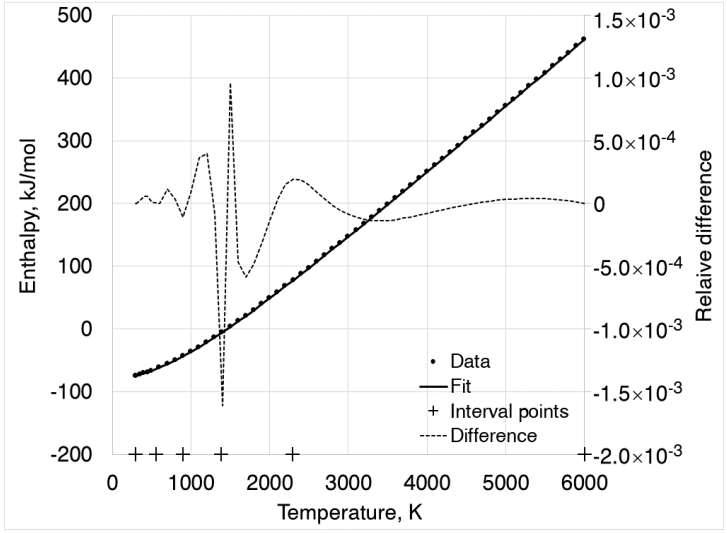


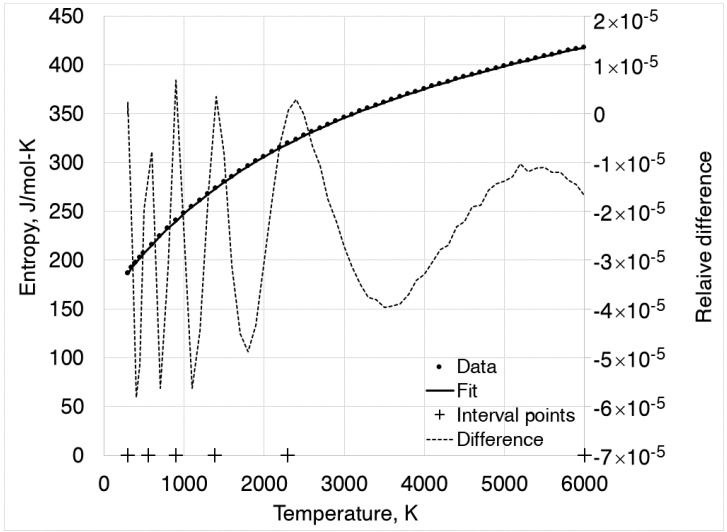


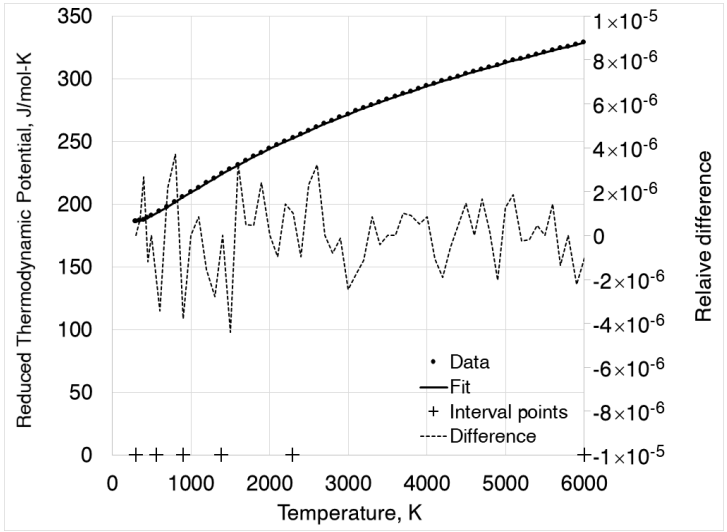
Methane (CH_4) – C1H4(g)

MW = 16.043 g/mol – $\Delta_f H^o$(298.15 K) = -74.873 kJ/mol – S^o(298.15 K) = 186.251 J/mol-K – p^o = 0.1 MPa



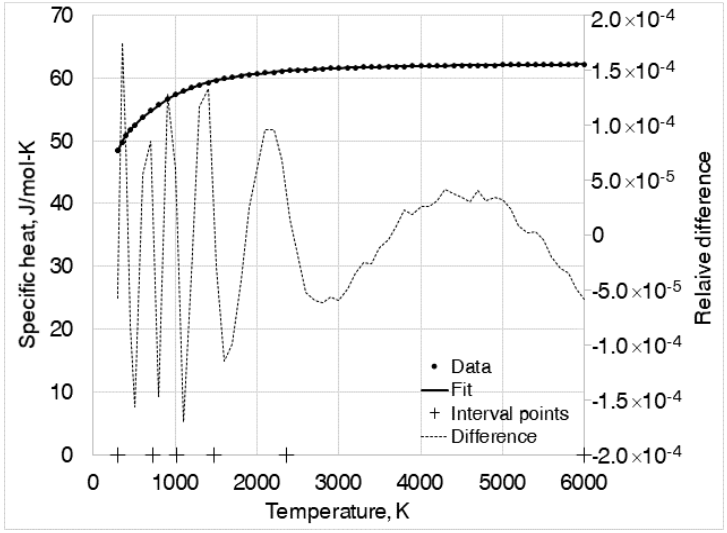


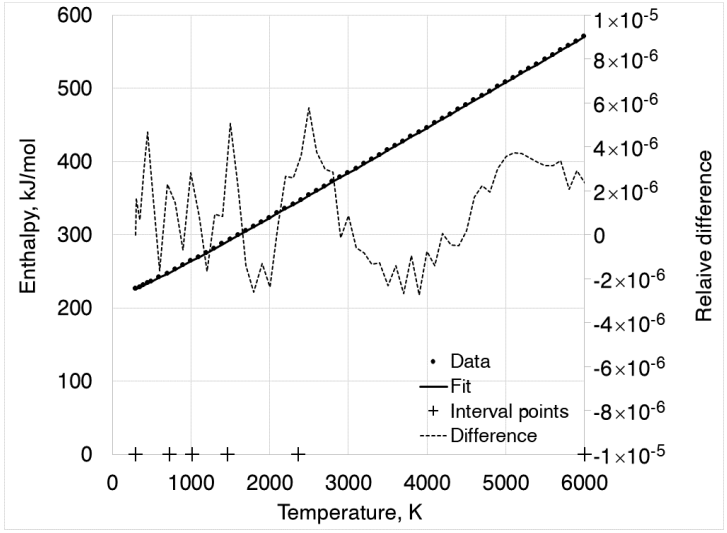


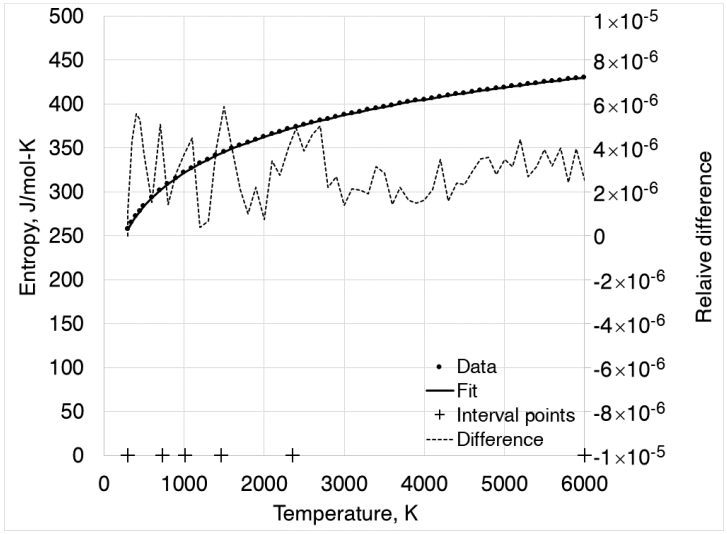


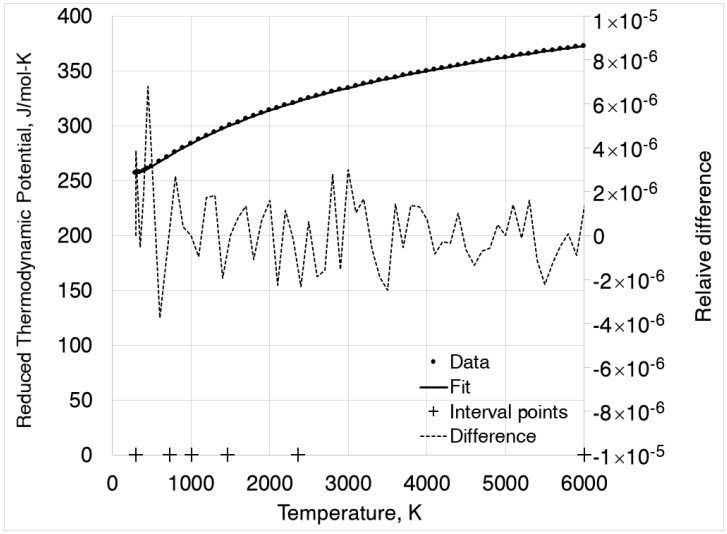
Cyanogen Iodide (CNI) – C1I1N1(g)

MW = 152.922 g/mol – $\Delta_f H^o$(298.15 K) = 225.936 kJ/mol – S^o(298.15 K) = 257.34 J/mol-K – p^o = 0.1 MPa



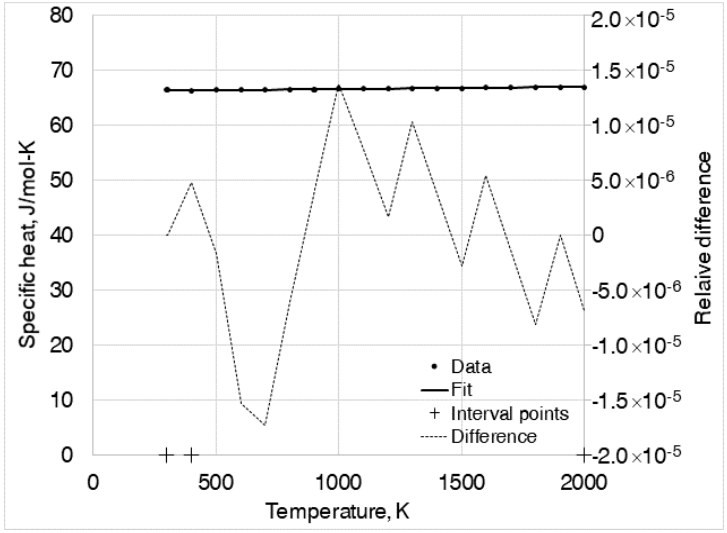


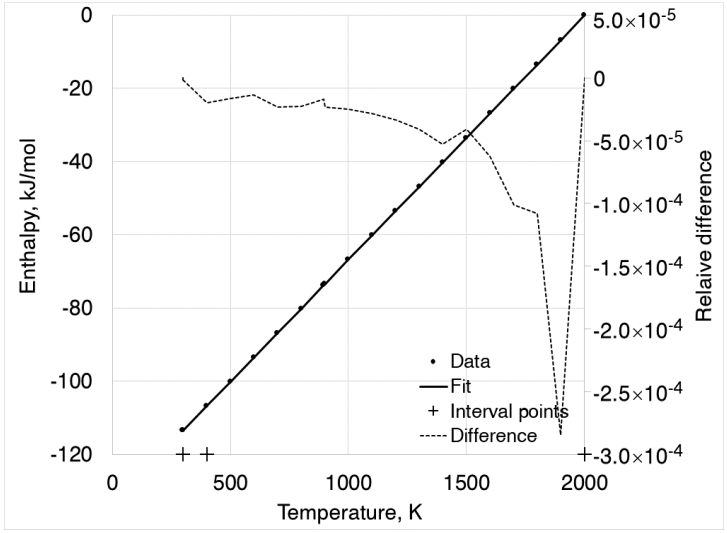


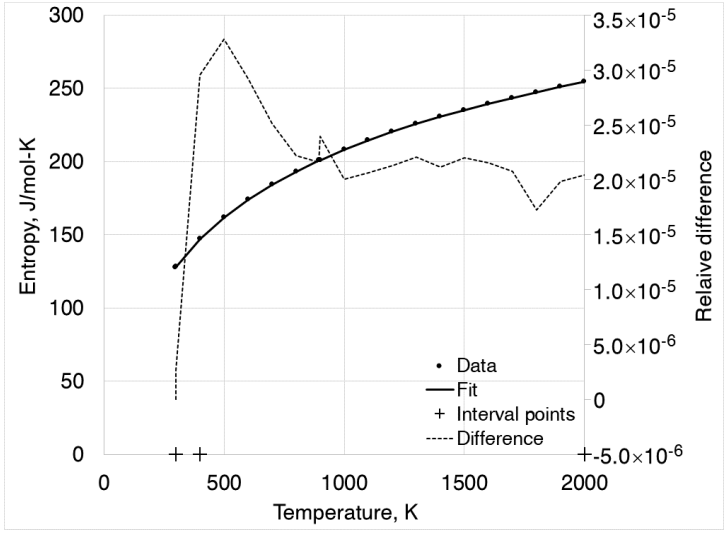


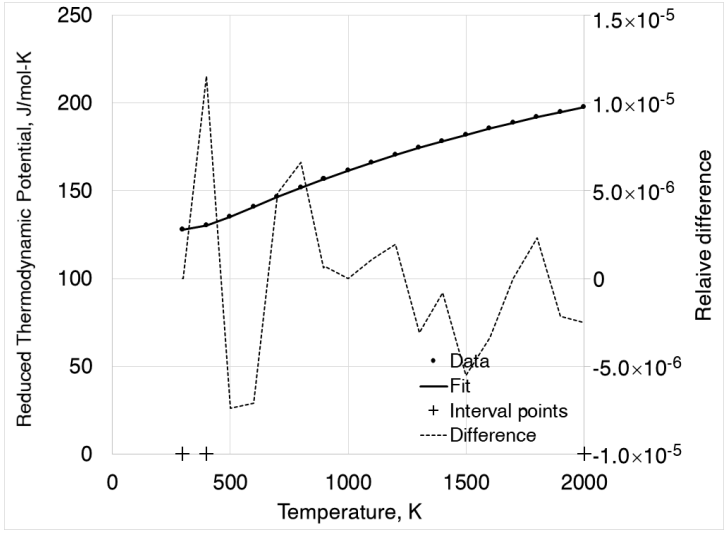
Potassium Cyanide (KCN) – C1K1N1(cr)

MW = 65.116 g/mol – $\Delta_f H^o$(298.15 K) = -113.47 kJ/mol – S^o(298.15 K) = 127.78 J/mol-K – p^o = 0.1 MPa



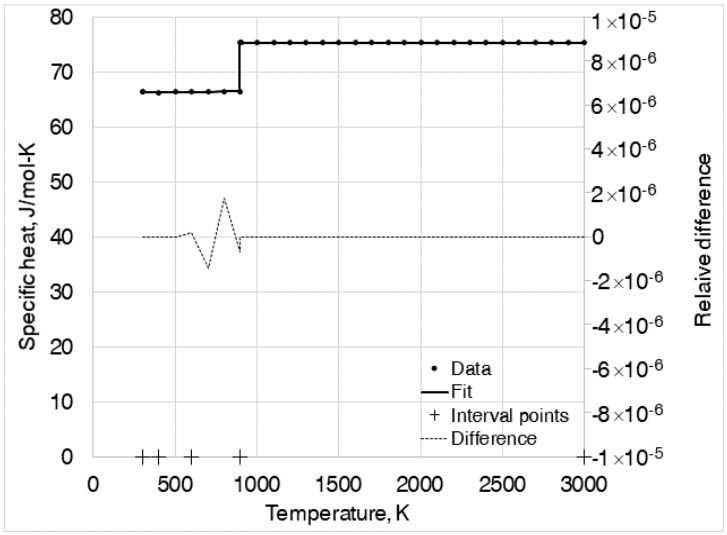


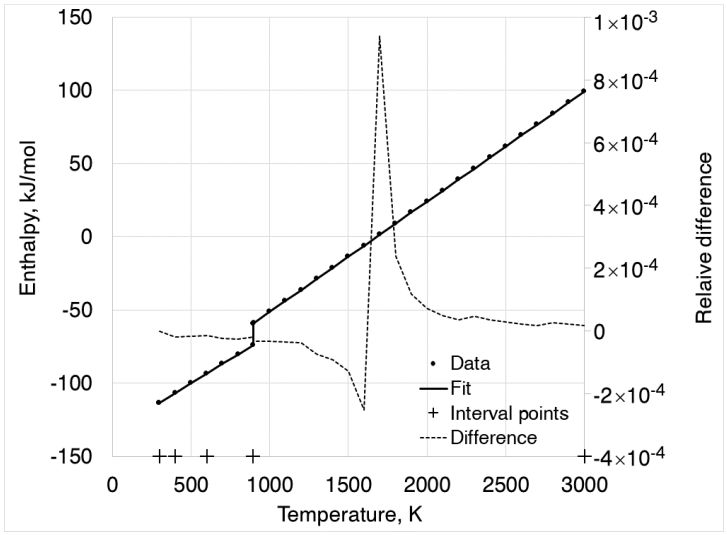


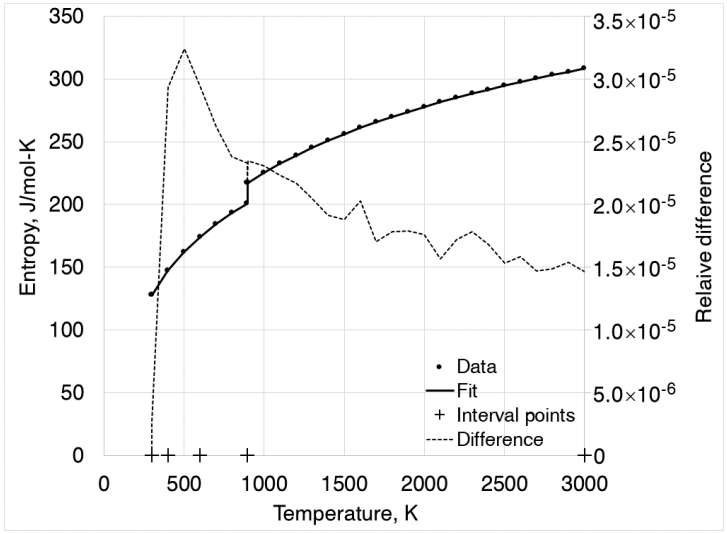


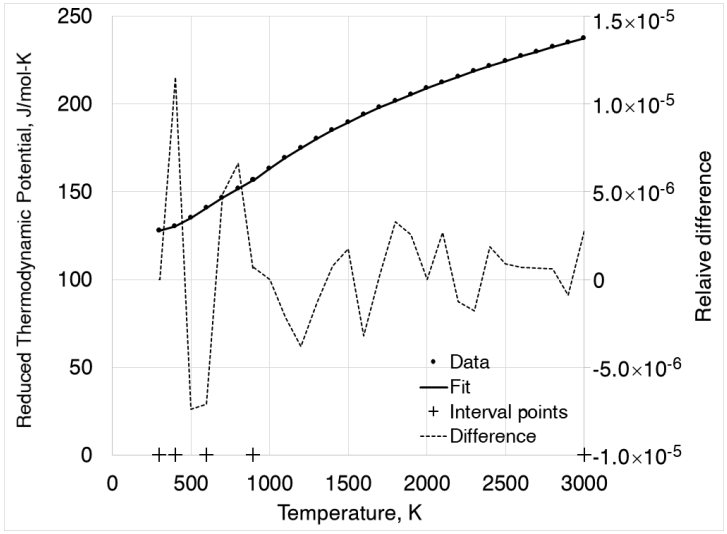
Potassium Cyanide (KCN) – C1K1N1(cr,l)

MW = 65.116 g/mol – $\Delta_f H^o$(298.15 K) = -113.47 kJ/mol – S^o(298.15 K) = 127.78 J/mol-K – p^o = 0.1 MPa



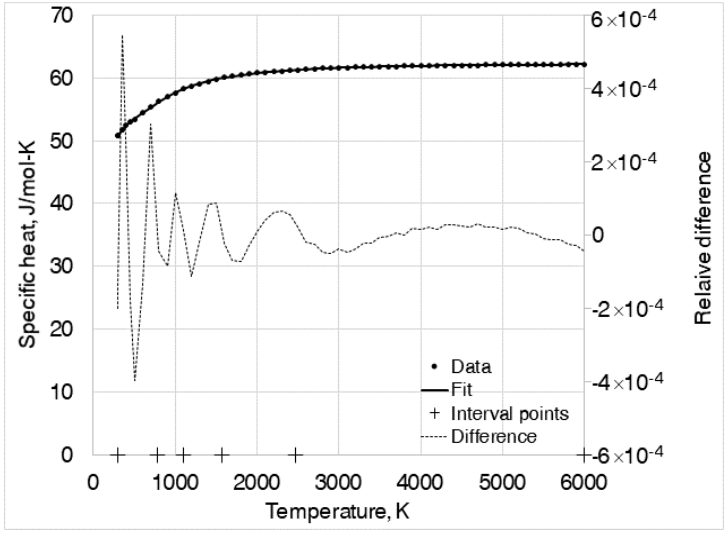


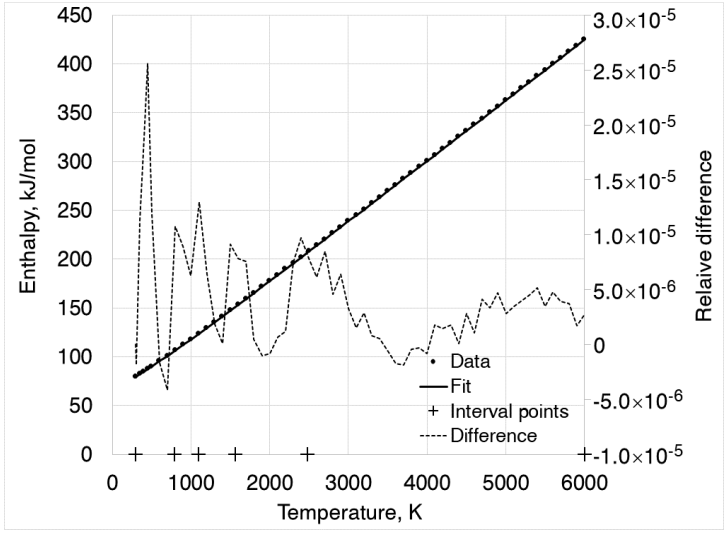


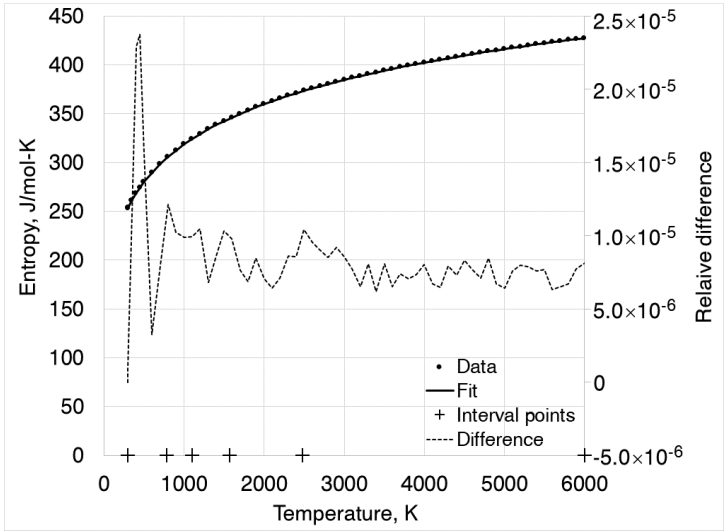


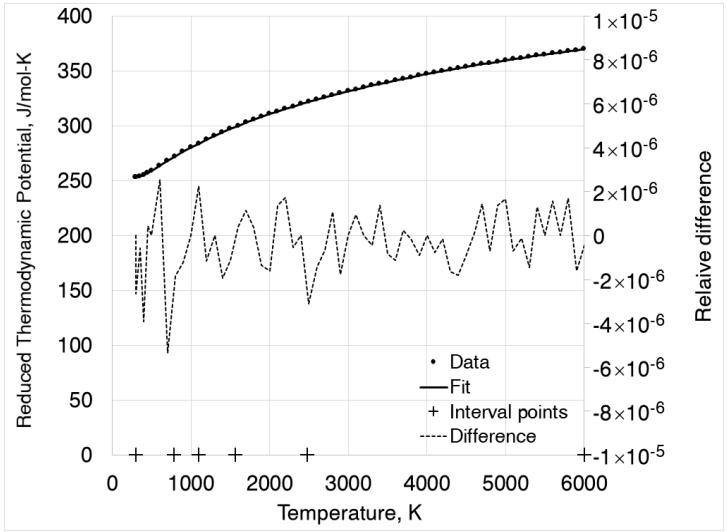
Potassium Cyanide (KCN) – C1K1N1(g)

MW = 65.116 g/mol – $\Delta_f H^o$(298.15 K) = 79.496 kJ/mol – S^o(298.15 K) = 253.146 J/mol-K – p^o = 0.1 MPa



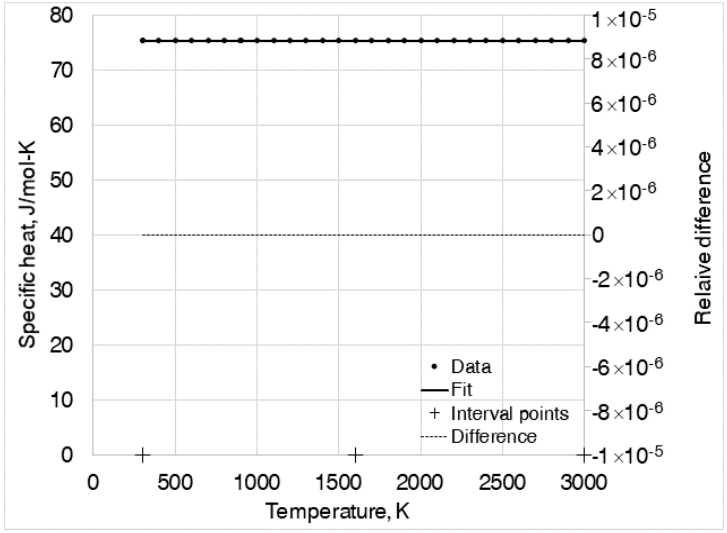


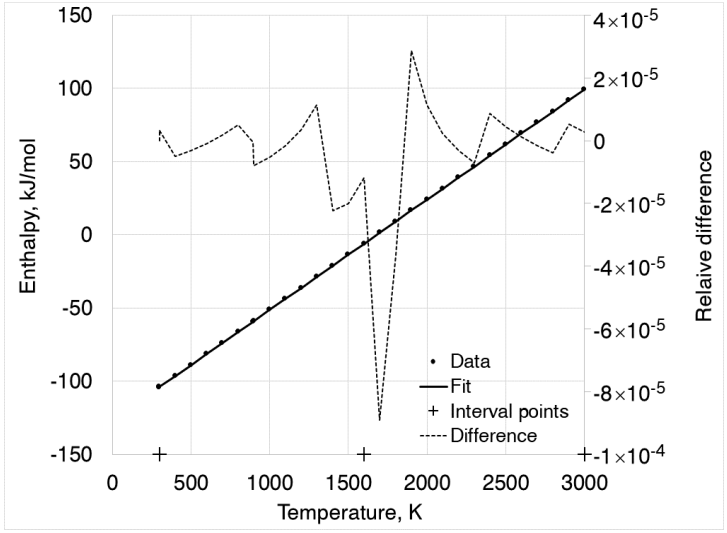


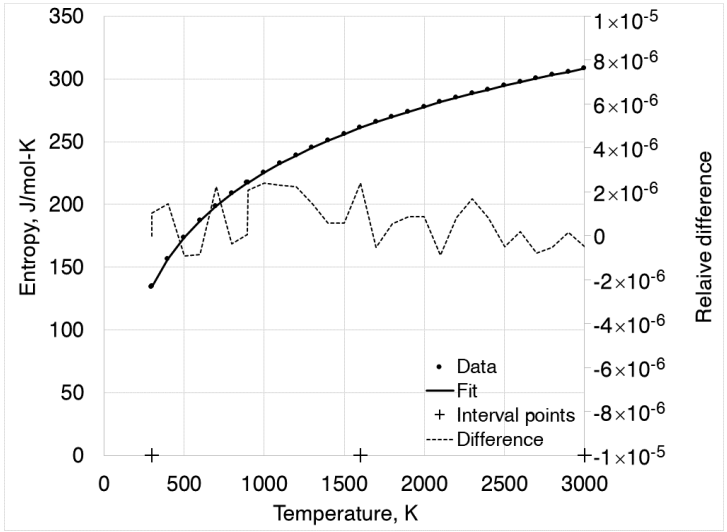


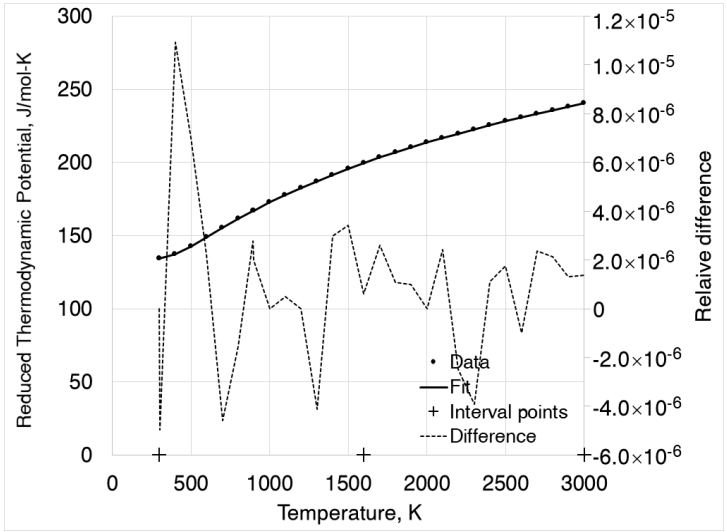
Potassium Cyanide (KCN) – C1K1N1(l)

MW = 65.116 g/mol – $\Delta_f H^o$(298.15 K) = -104.148 kJ/mol – S^o(298.15 K) = 134.321 J/mol-K – p^o = 0.1 MPa



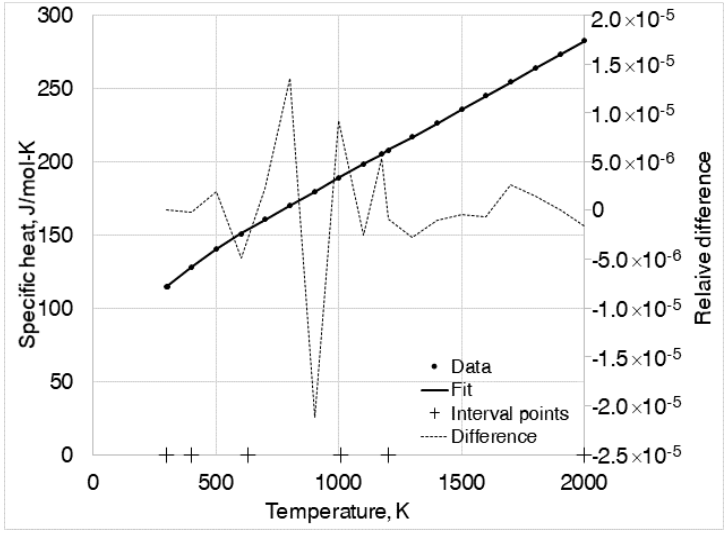


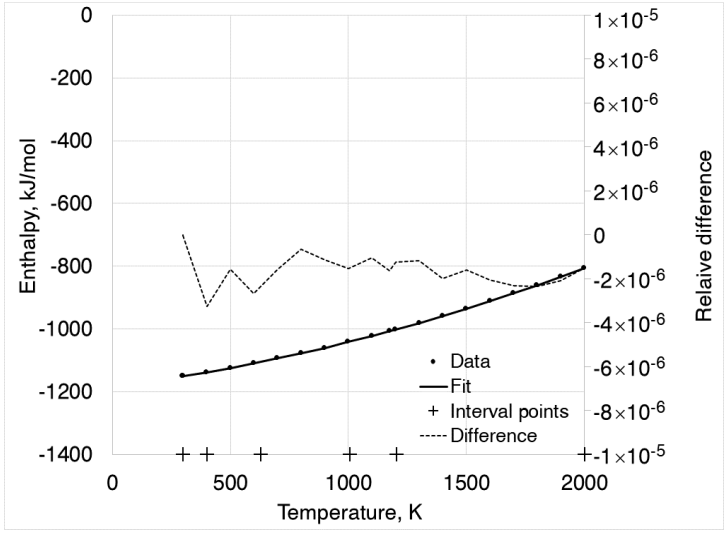


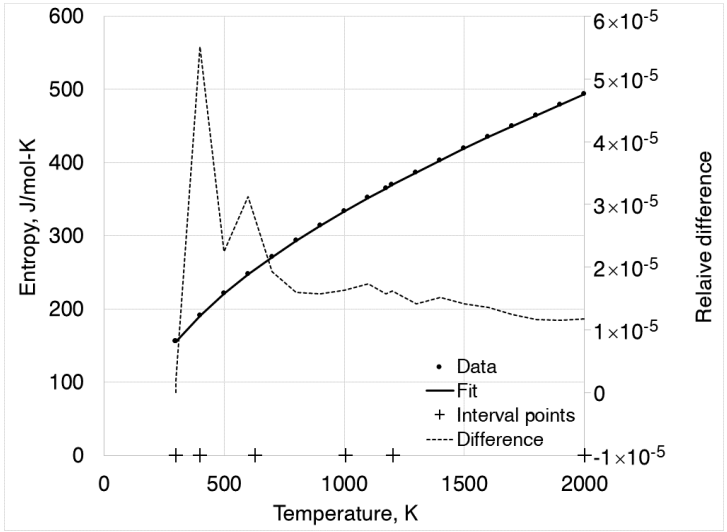


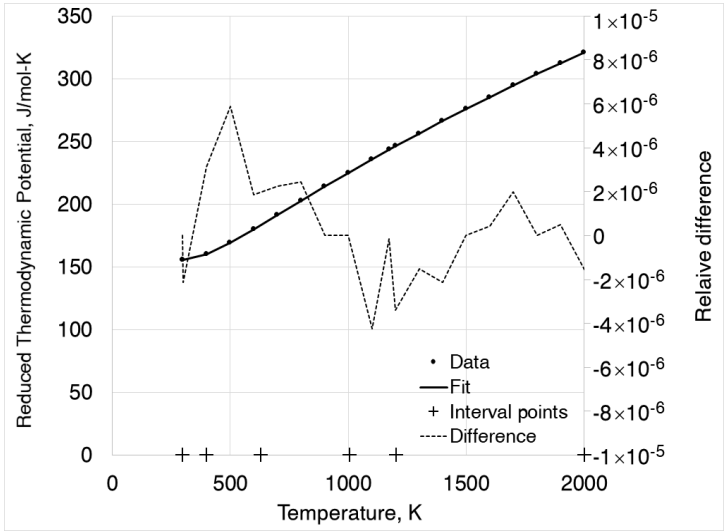
Potassium Carbonate (K_2CO_3) – C1K2O3(cr)

MW = 138.206 g/mol – $\Delta_f H^o$(298.15 K) = -1150.182 kJ/mol – S^o(298.15 K) = 155.519 J/mol-K – p^o = 0.1 MPa



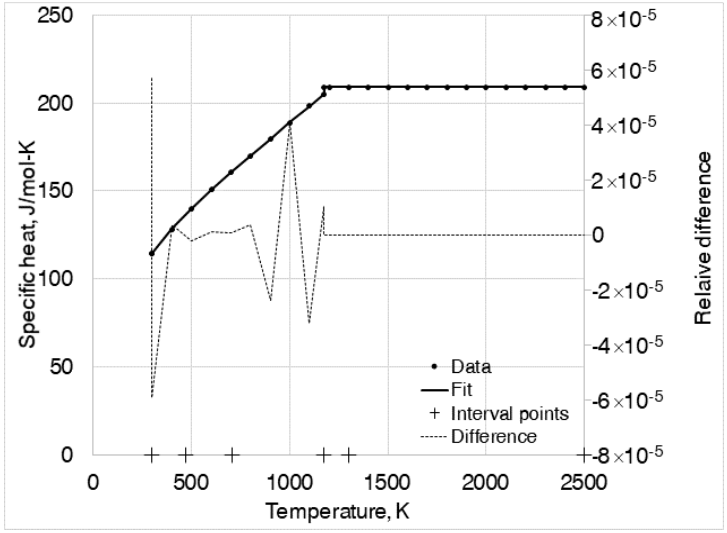


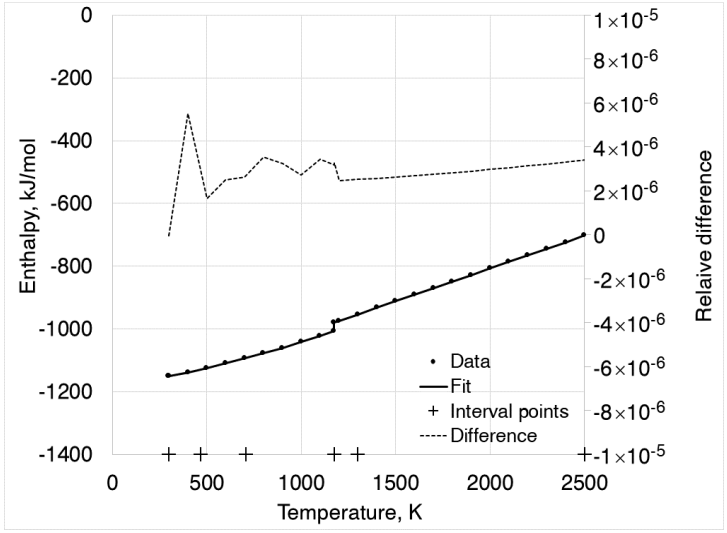


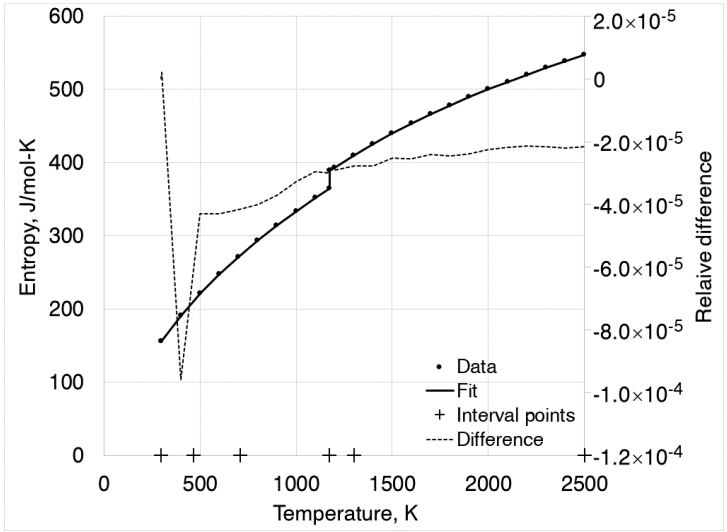


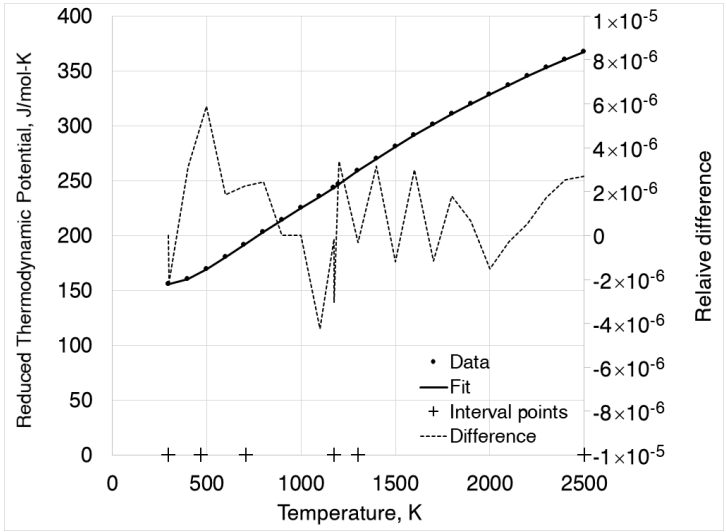
Potassium Carbonate (K_2CO_3) – C1K2O3(cr,l)

MW = 138.206 g/mol – $\Delta_f H^o$(298.15 K) = -1150.182 kJ/mol – S^o(298.15 K) = 155.519 J/mol-K – p^o = 0.1 MPa



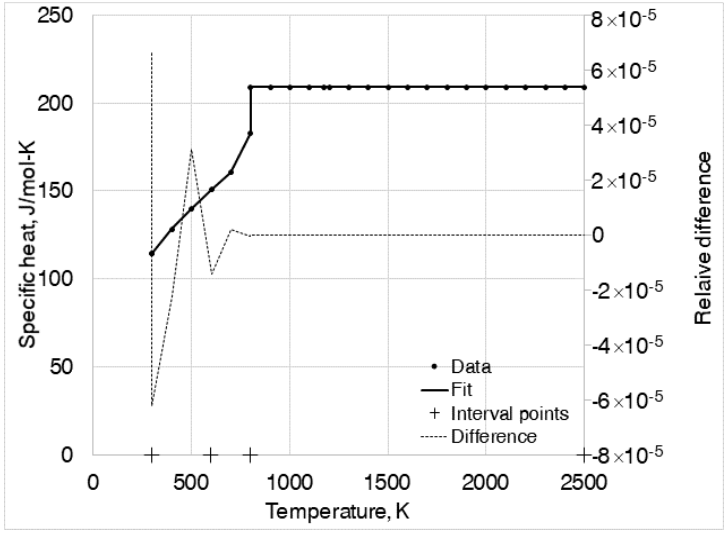


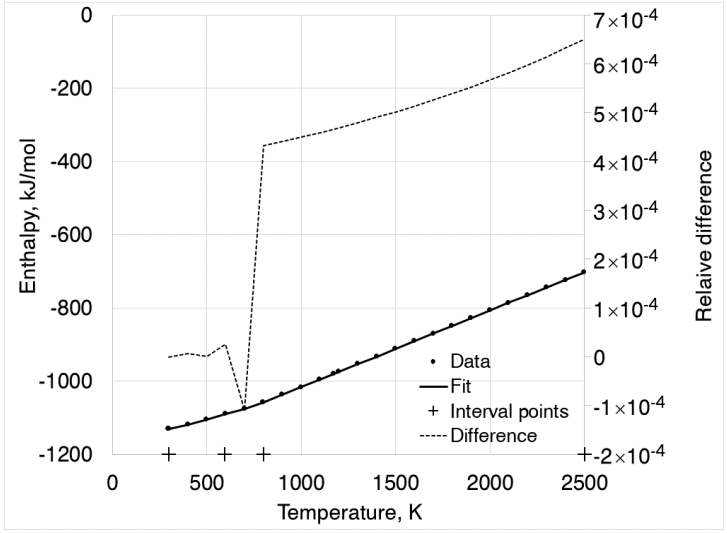


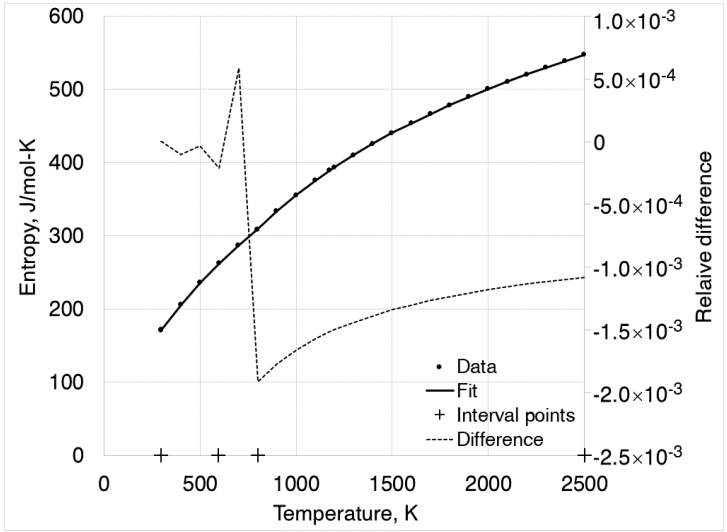


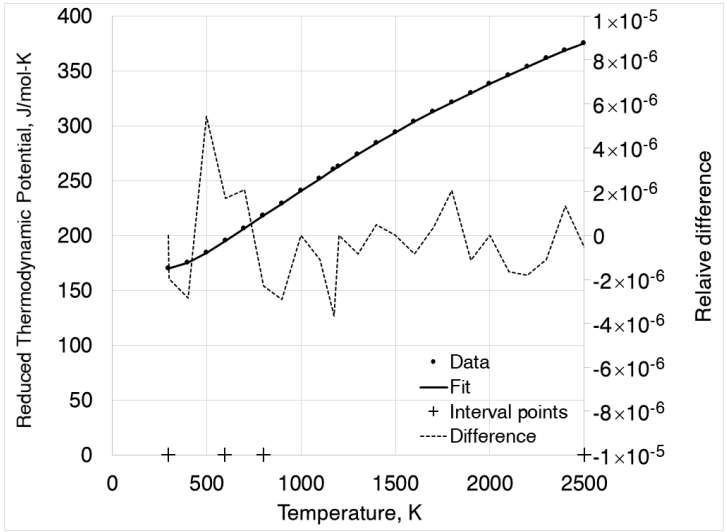
Potassium Carbonate (K_2CO_3) – C1K2O3(l)

MW = 138.206 g/mol – $\Delta_f H^o$(298.15 K) = -1130.609 kJ/mol – S^o(298.15 K) = 170.373 J/mol-K – p^o = 0.1 MPa



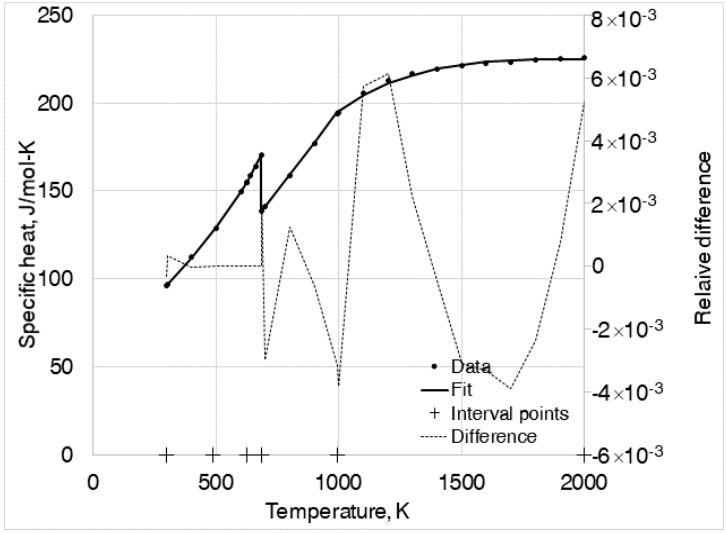


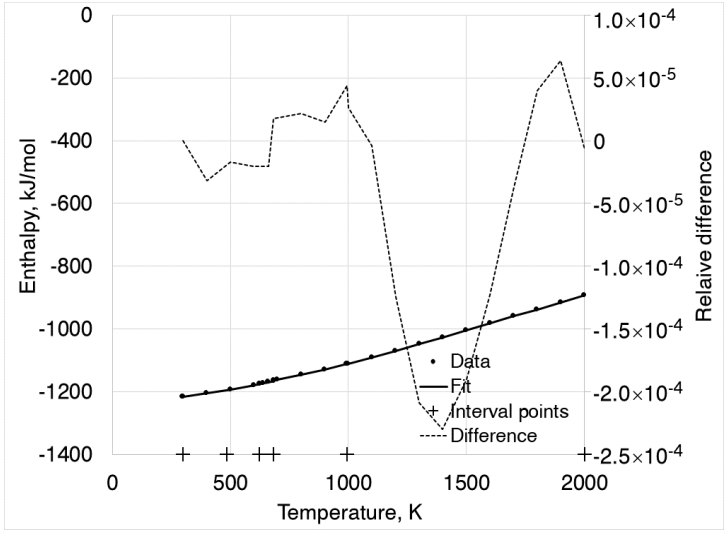


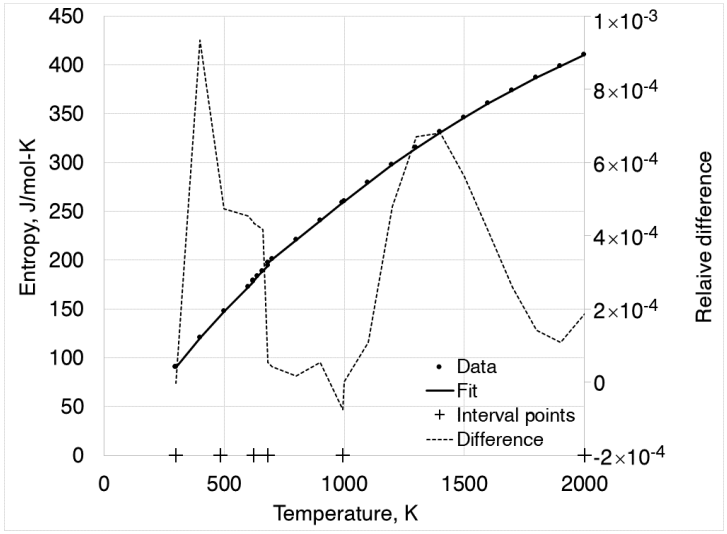


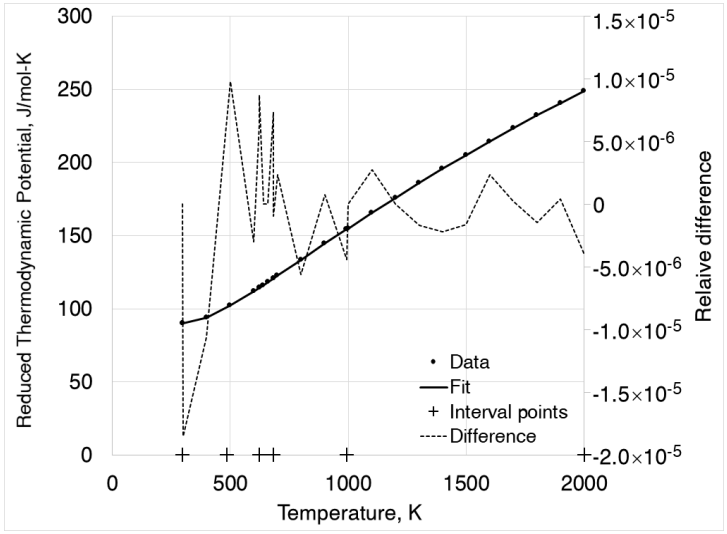
Lithium Carbonate (Li_2CO_3) – C1Li2O3(cr)

MW = 73.891 g/mol – $\Delta_f H^o$(298.15 K) = -1216.038 kJ/mol – S^o(298.15 K) = 90.171 J/mol-K – p^o = 0.1 MPa



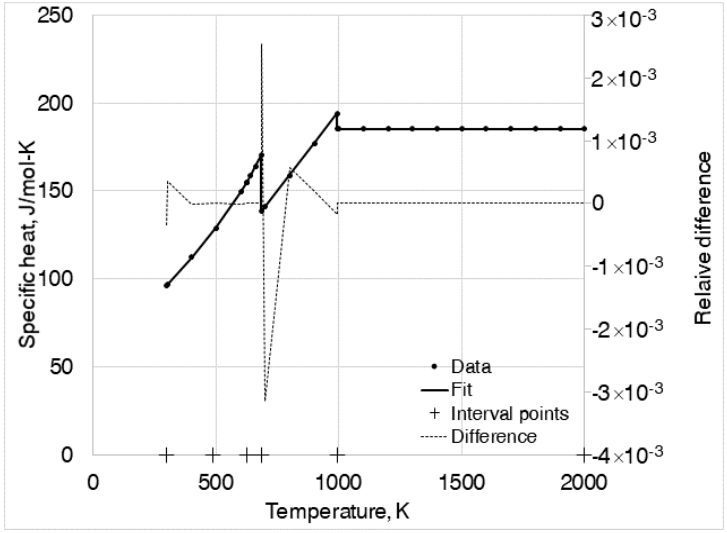


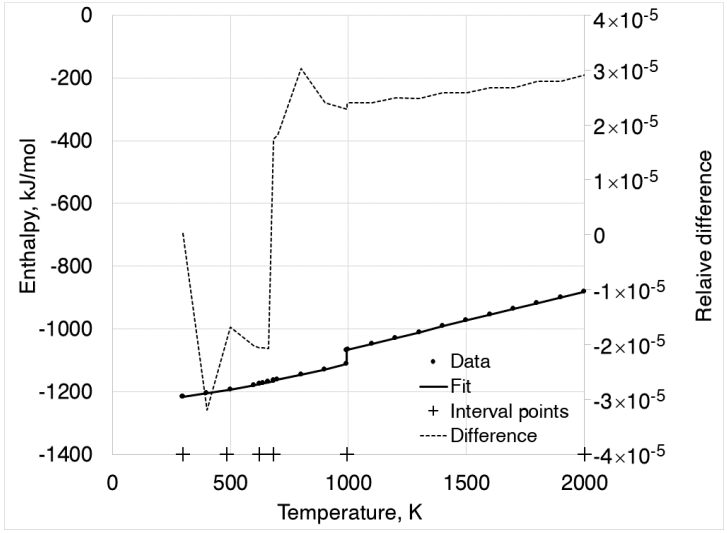


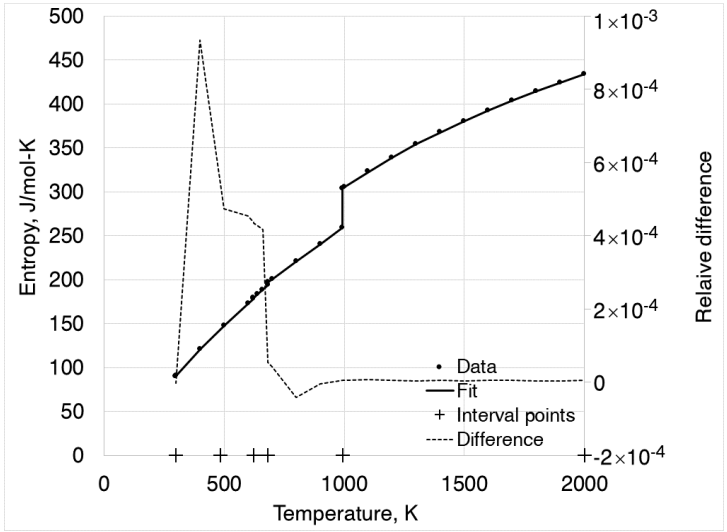


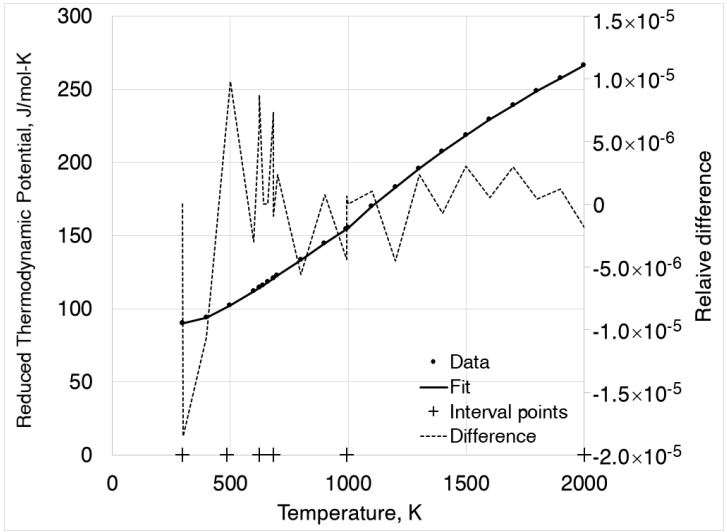
Lithium Carbonate (Li_2CO_3) – C1Li2O3(cr,l)

MW = 73.891 g/mol – $\Delta_f H^o$(298.15 K) = -1216.038 kJ/mol – S^o(298.15 K) = 90.171 J/mol-K – p^o = 0.1 MPa



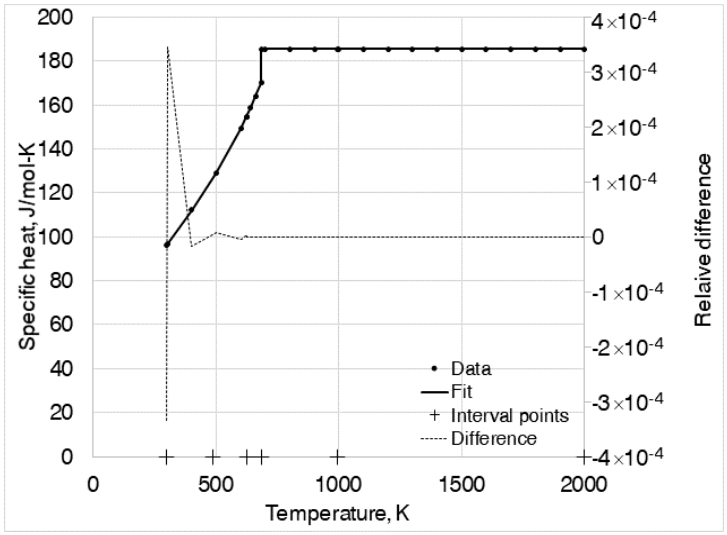


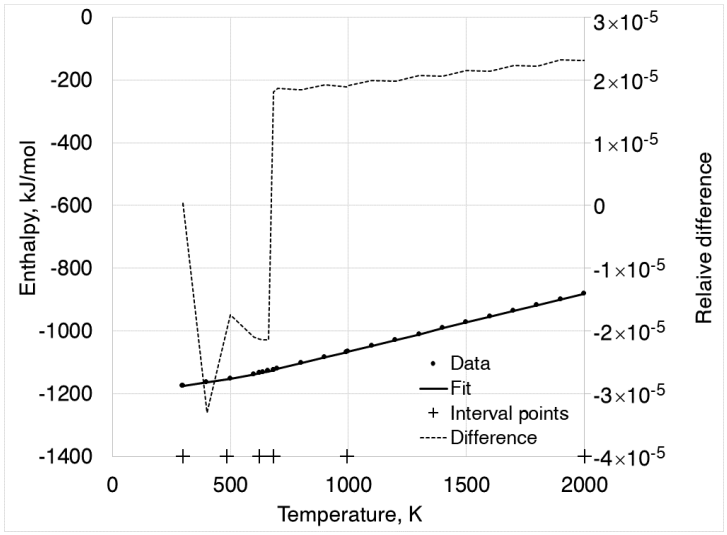


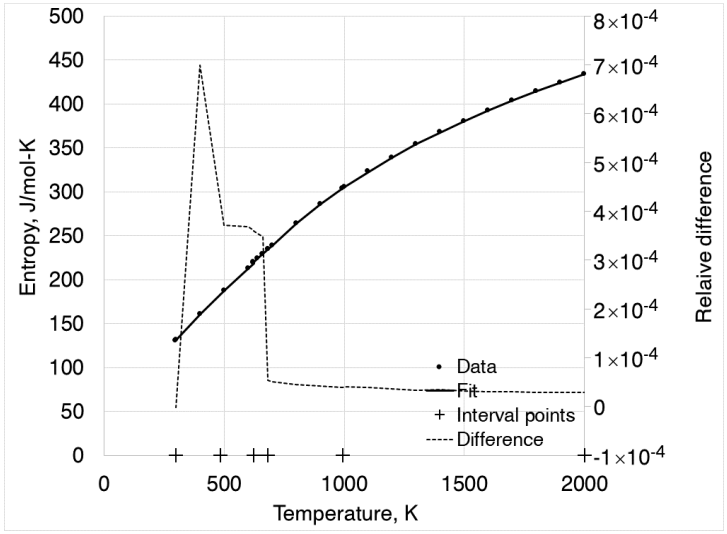


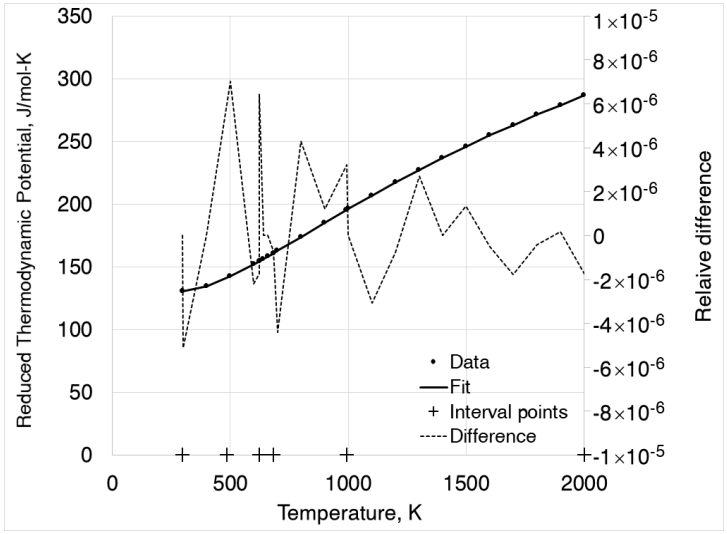
Lithium Carbonate (Li_2CO_3) – C1Li2O3(l)

MW = 73.891 g/mol – $\Delta_f H^o$(298.15 K) = -1175.083 kJ/mol – S^o(298.15 K) = 130.567 J/mol-K – p^o = 0.1 MPa



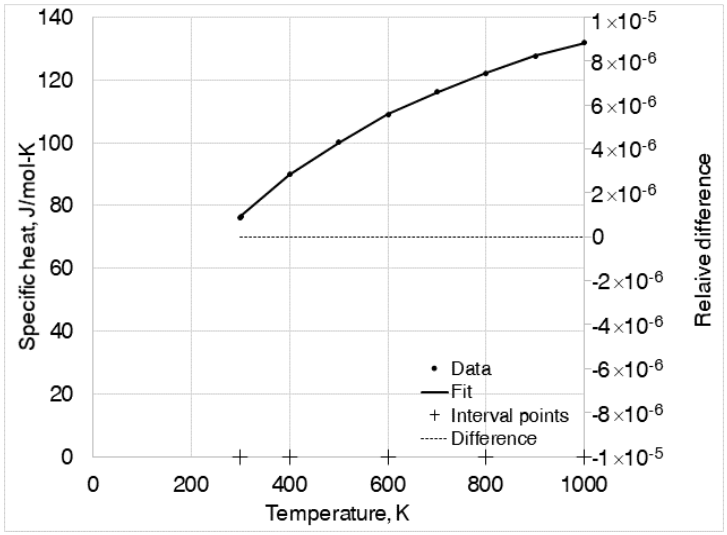


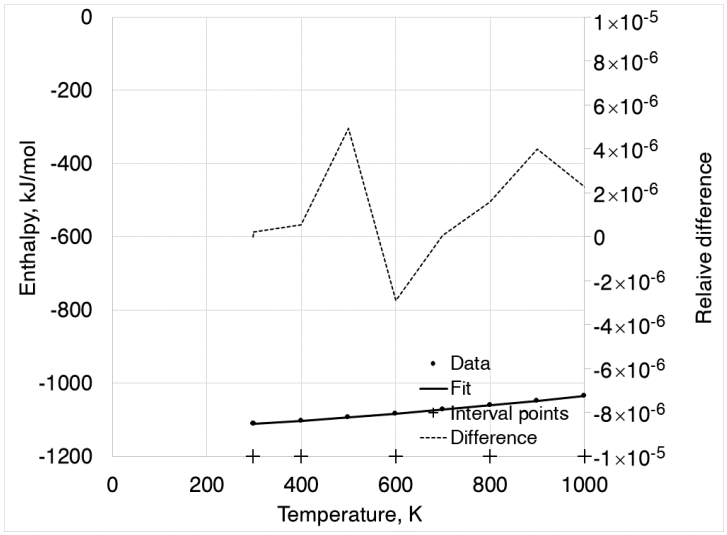


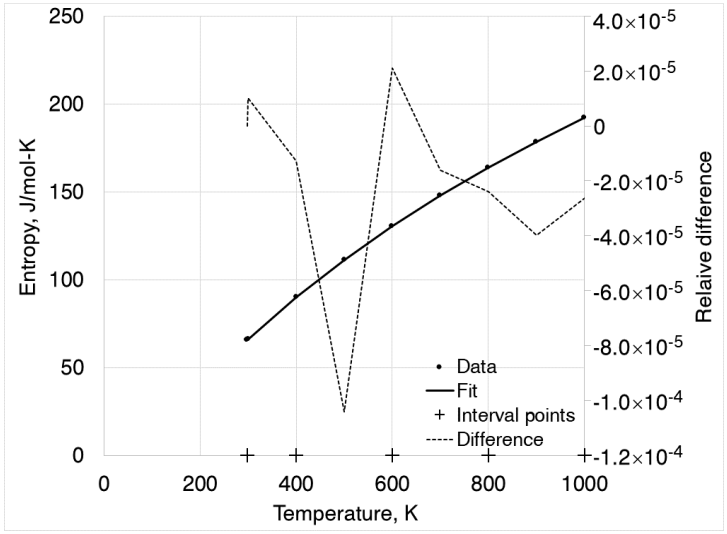


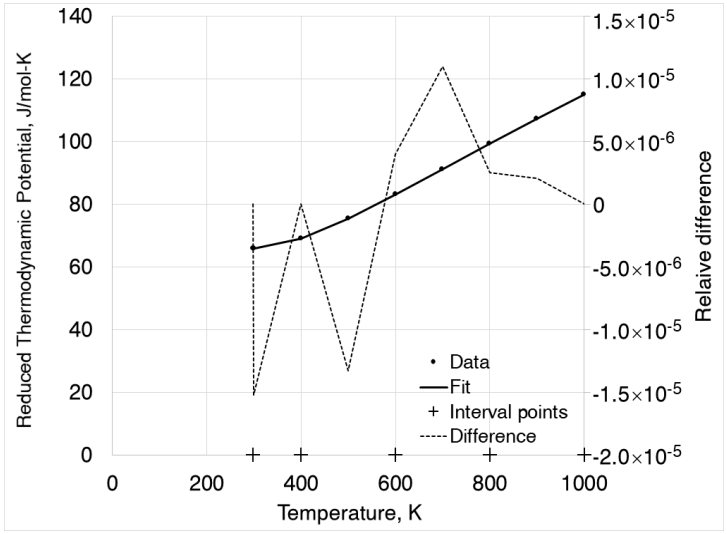
Magnesium Carbonate ($MgCO_3$) – C1Mg1O3(cr)

MW = 84.314 g/mol – $\Delta_f H^o$(298.15 K) = -1111.689 kJ/mol – S^o(298.15 K) = 65.854 J/mol-K – p^o = 0.1 MPa



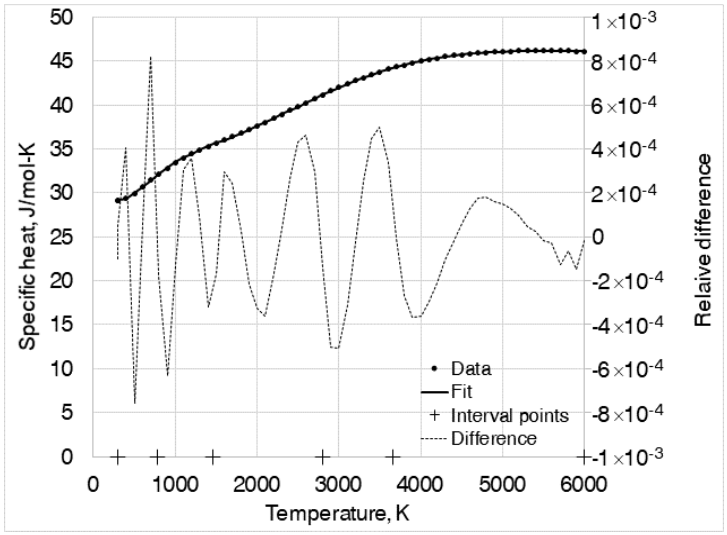


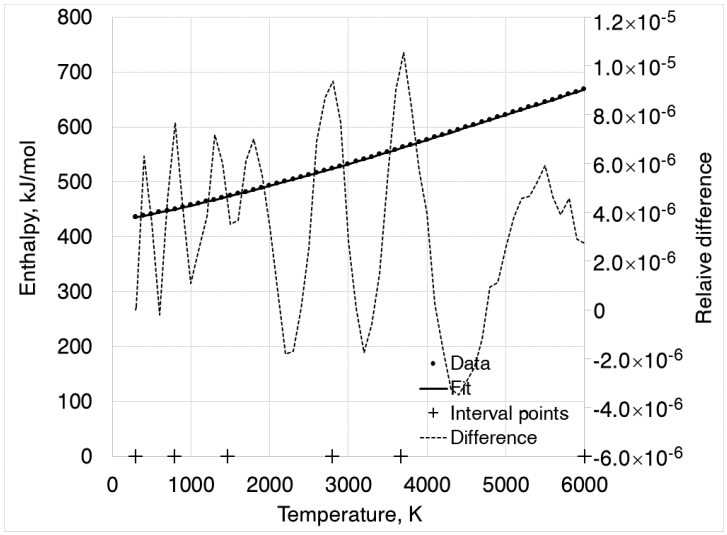


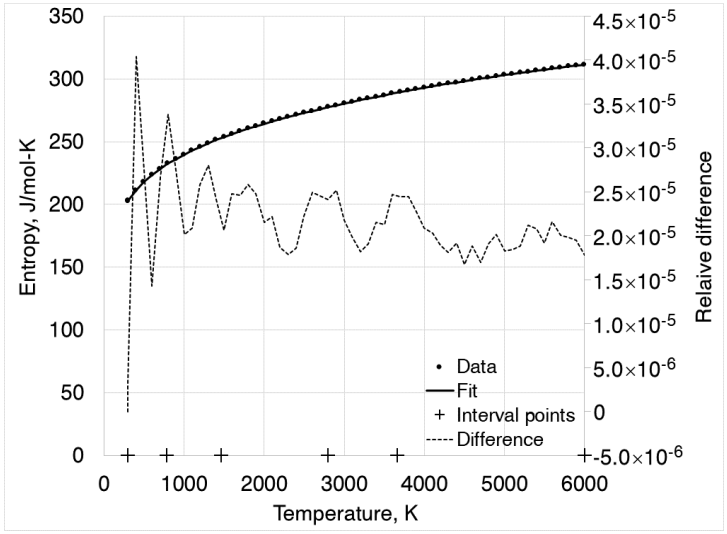


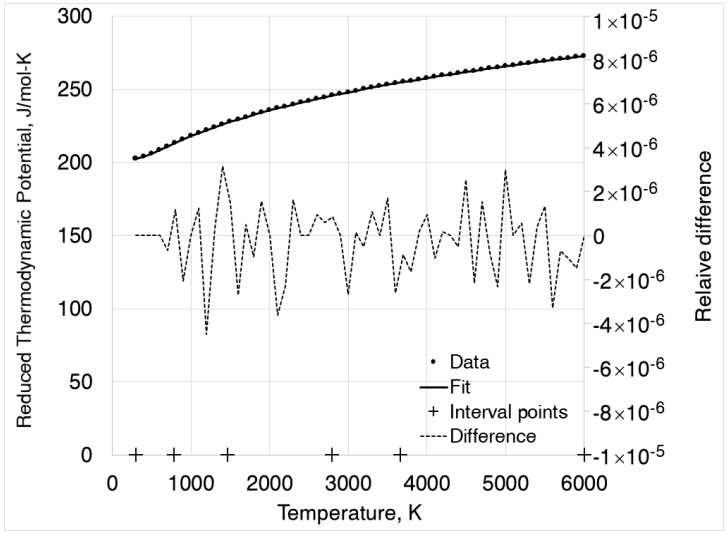
Cyanogen (CN) – C1N1(g)

MW = 26.018 g/mol – $\Delta_f H^o$(298.15 K) = 435.136 kJ/mol – S^o(298.15 K) = 202.638 J/mol-K – p^o = 0.1 MPa



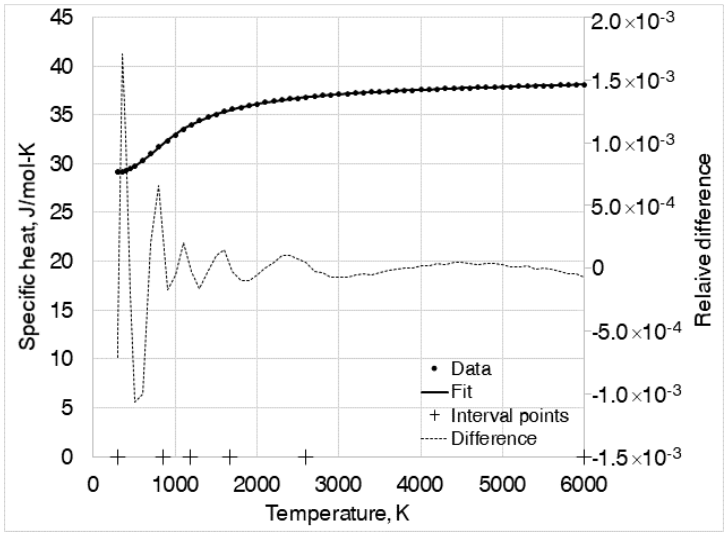


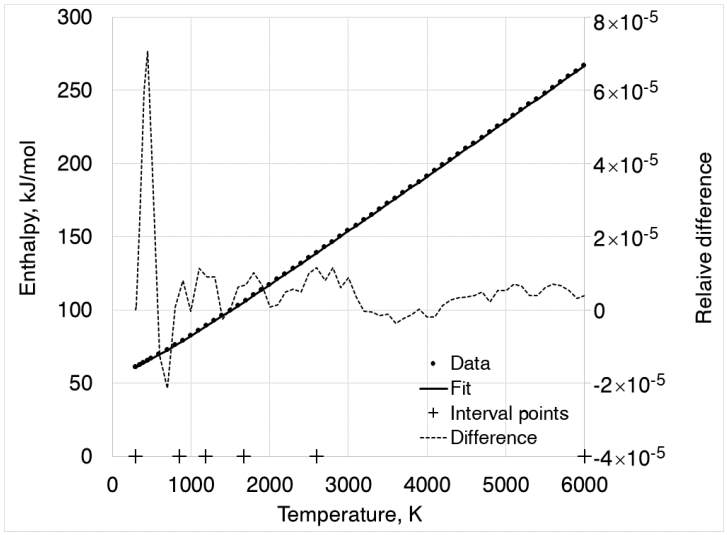


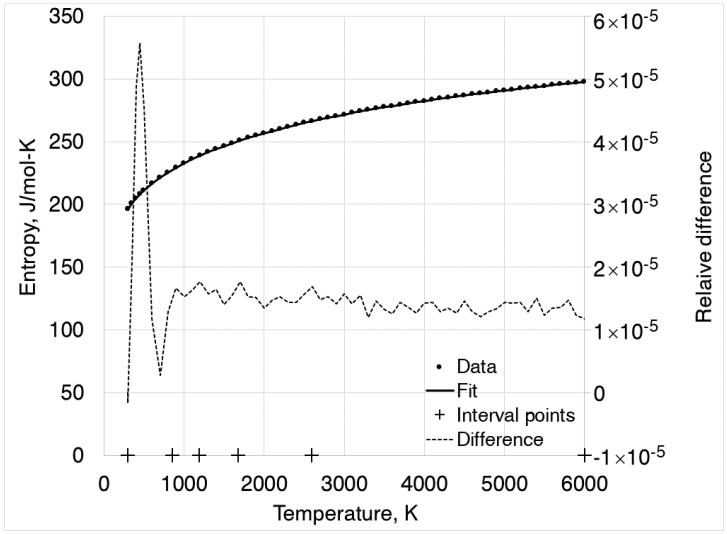


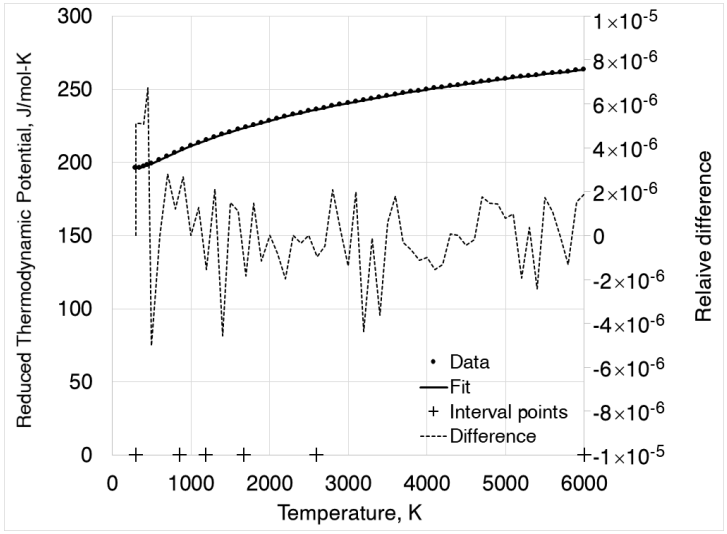
Cyanide (CN^-) – C1N1-(g)

MW = 26.018 g/mol – $\Delta_f H^o$(298.15 K) = 60.668 kJ/mol – S^o(298.15 K) = 195.976 J/mol-K – p^o = 0.1 MPa



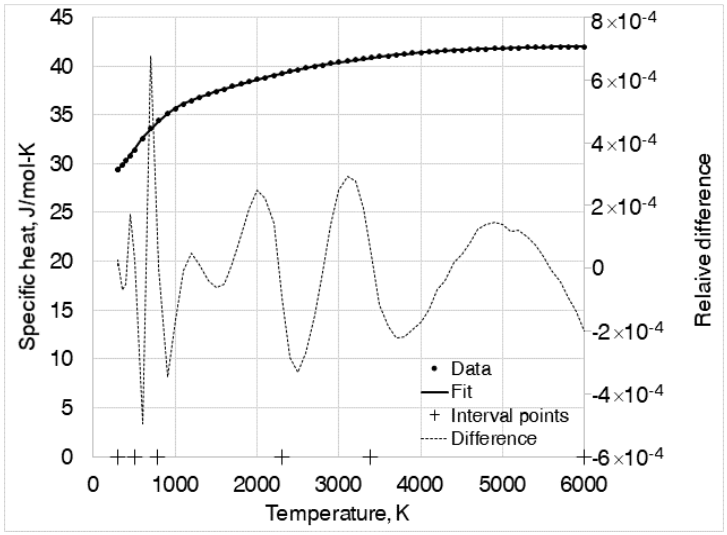


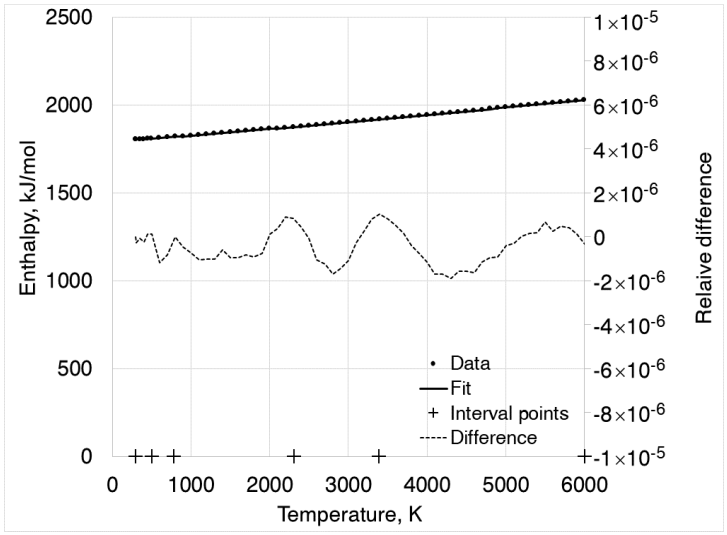


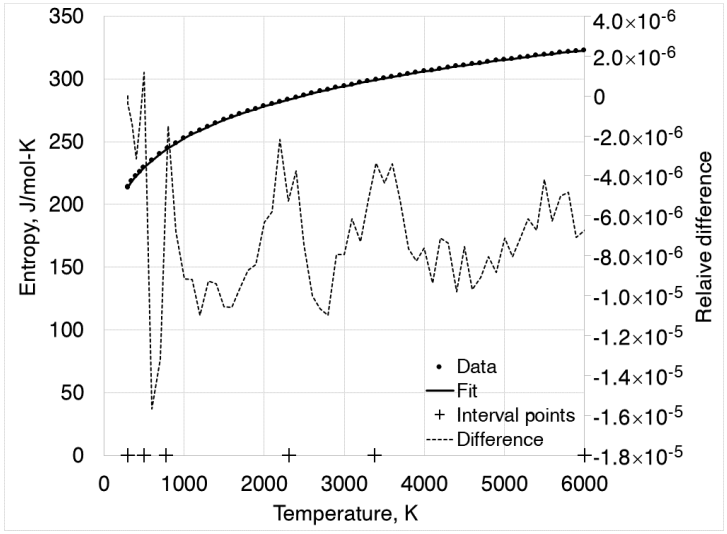


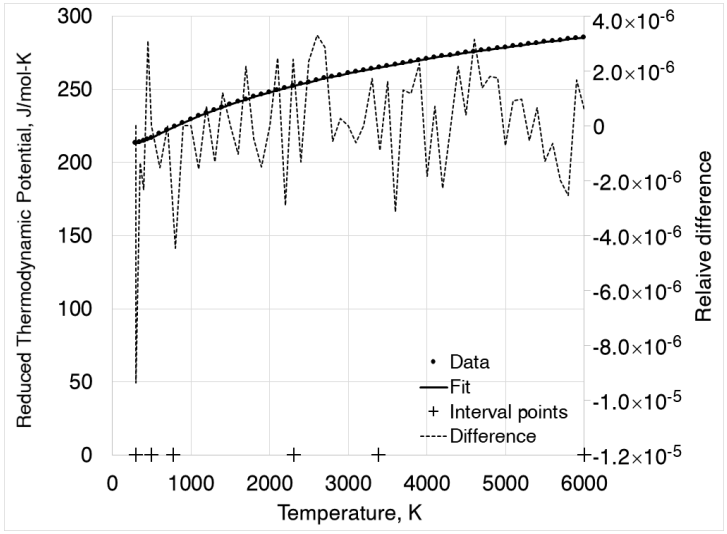
Cyanogen, Ion (CN^+) – C1N1+(g)

MW = 26.017 g/mol – $\Delta_f H^o$(298.15 K) = 1802.76 kJ/mol – S^o(298.15 K) = 213.437 J/mol-K – p^o = 0.1 MPa



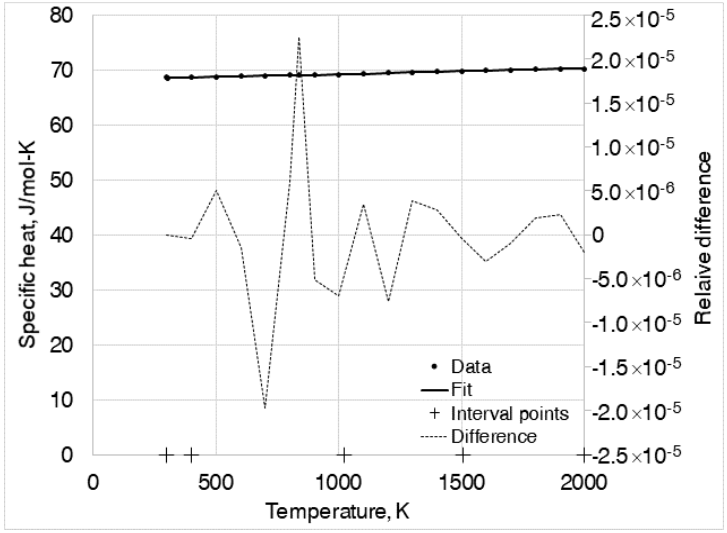


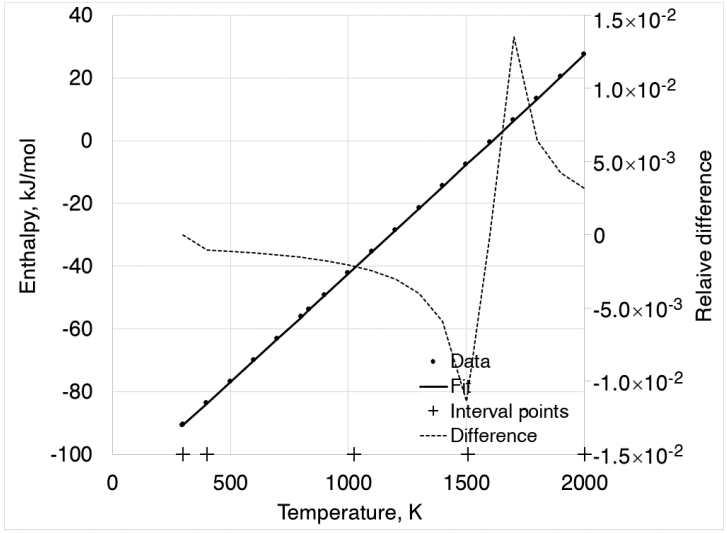


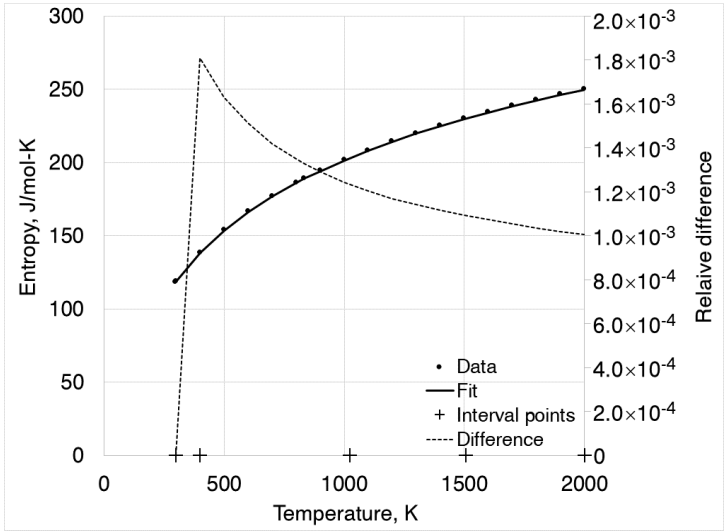


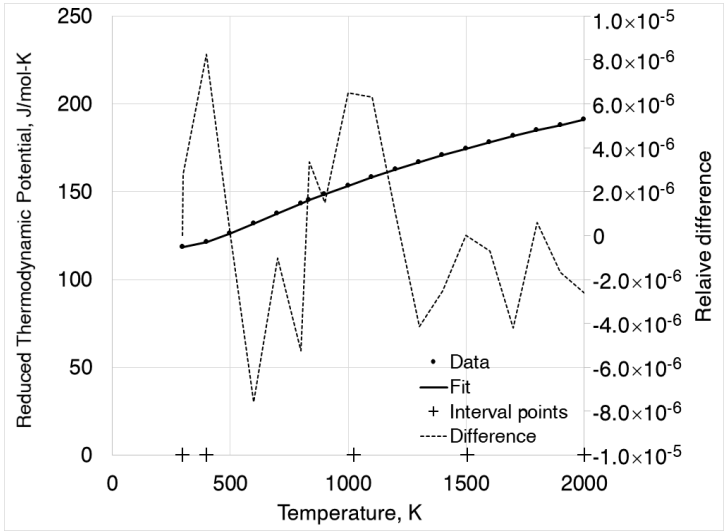
Sodium Cyanide (NaCN) – C1N1Na1(cr)

MW = 49.007 g/mol – $\Delta_f H^o$(298.15 K) = -90.709 kJ/mol – S^o(298.15 K) = 118.467 J/mol-K – p^o = 0.1 MPa



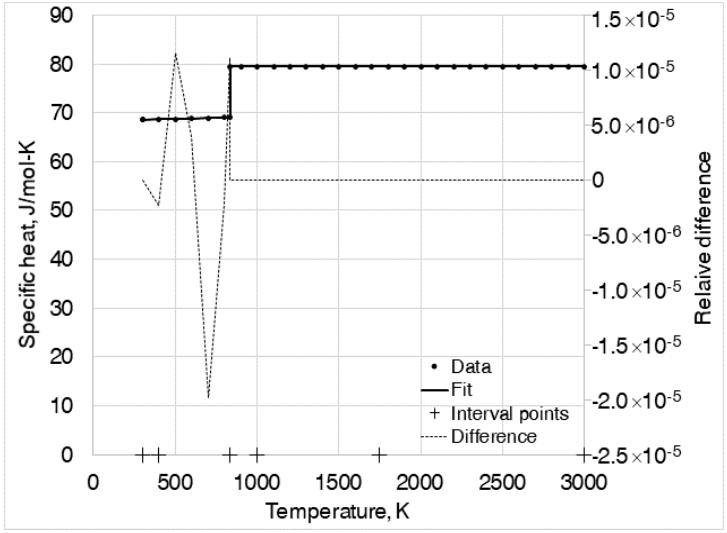


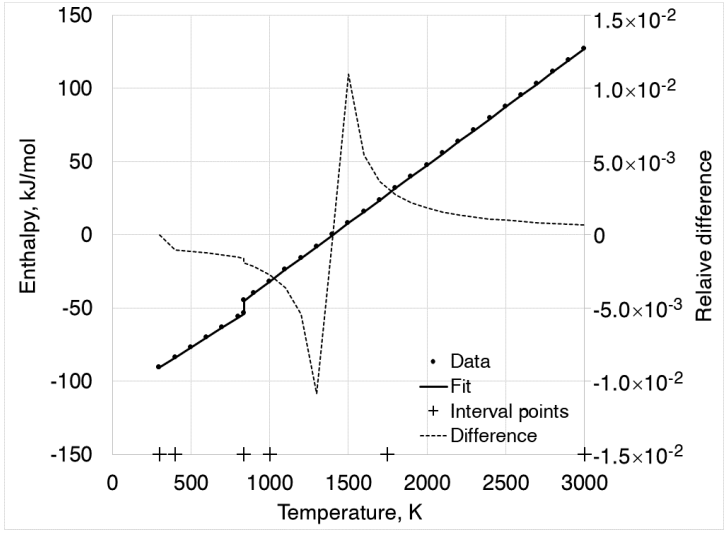


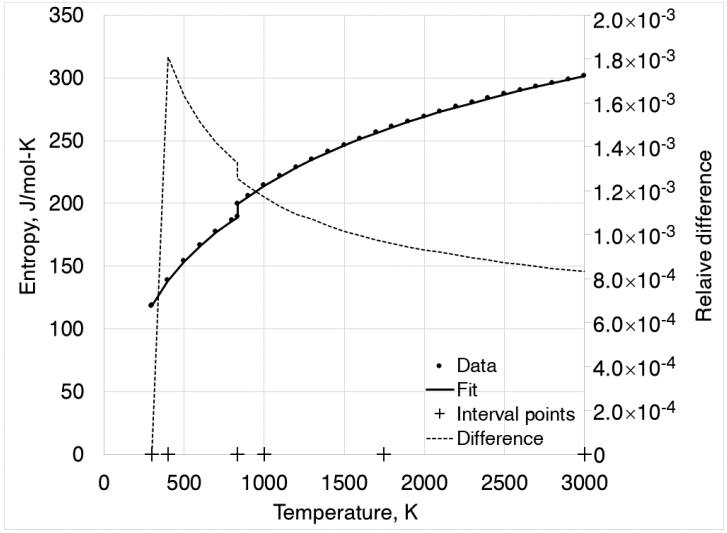


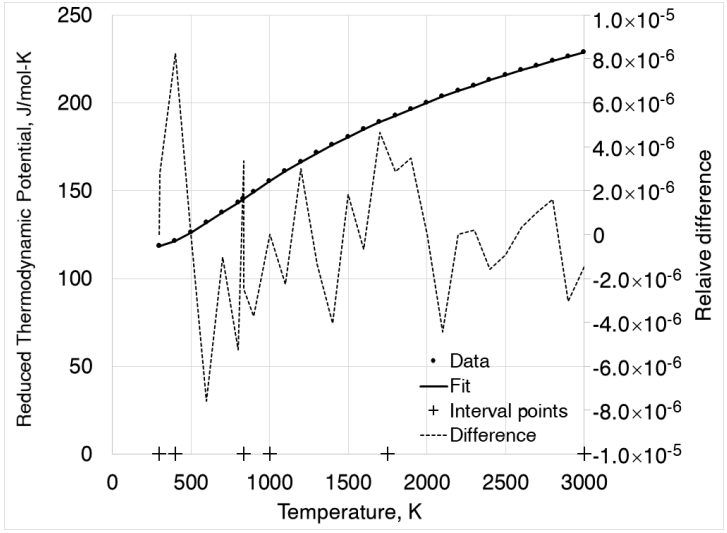
Sodium Cyanide (NaCN) – C1N1Na1(cr,l)

MW = 49.007 g/mol – $\Delta_f H^o$(298.15 K) = -90.709 kJ/mol – S^o(298.15 K) = 118.467 J/mol-K – p^o = 0.1 MPa



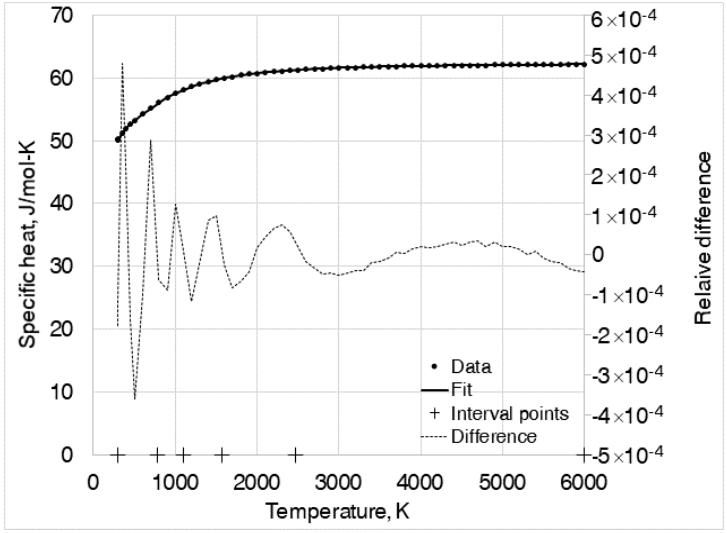


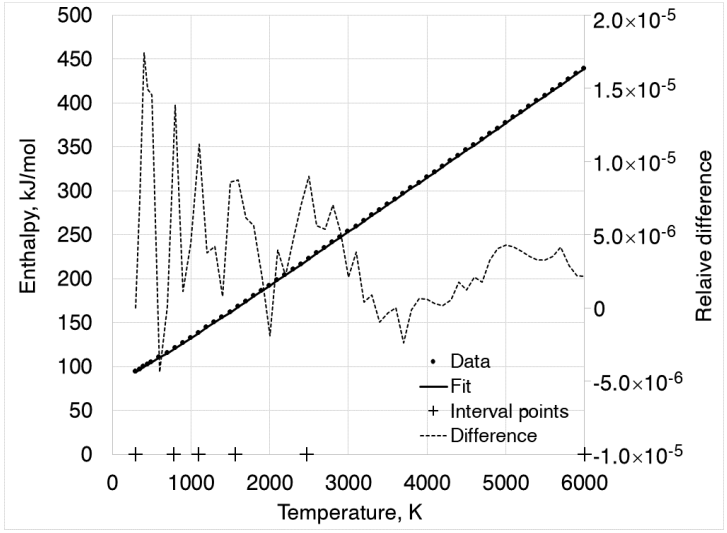


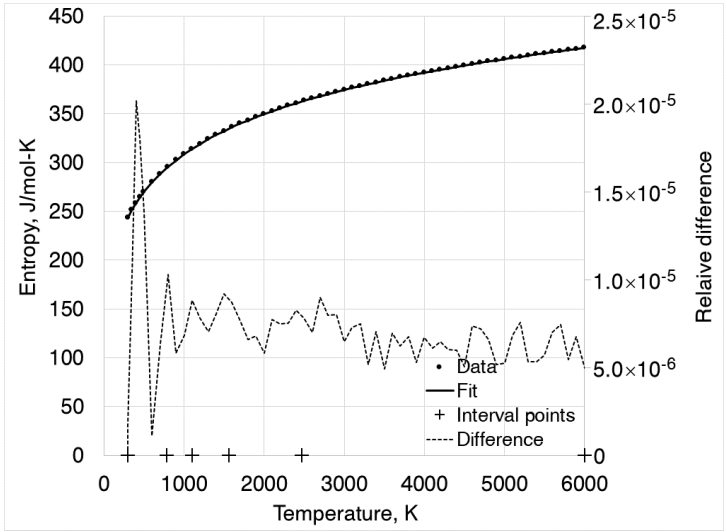


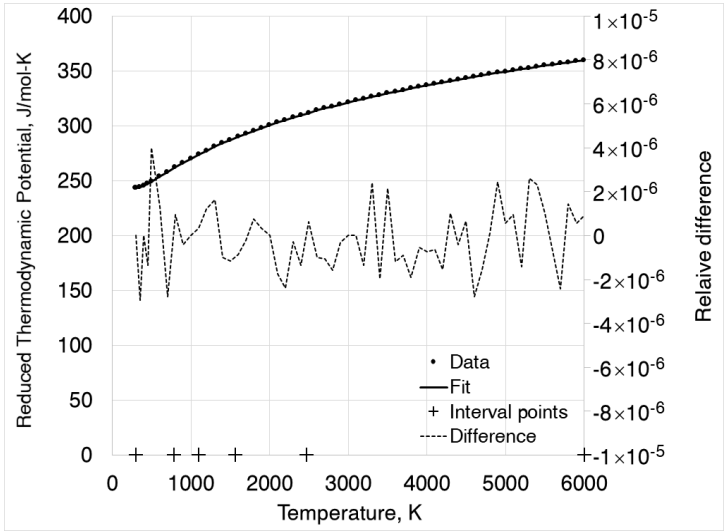
Sodium Cyanide (NaCN) – C1N1Na1(g)

MW = 49.007 g/mol – $\Delta_f H^o$(298.15 K) = 94.266 kJ/mol – S^o(298.15 K) = 243.367 J/mol-K – p^o = 0.1 MPa



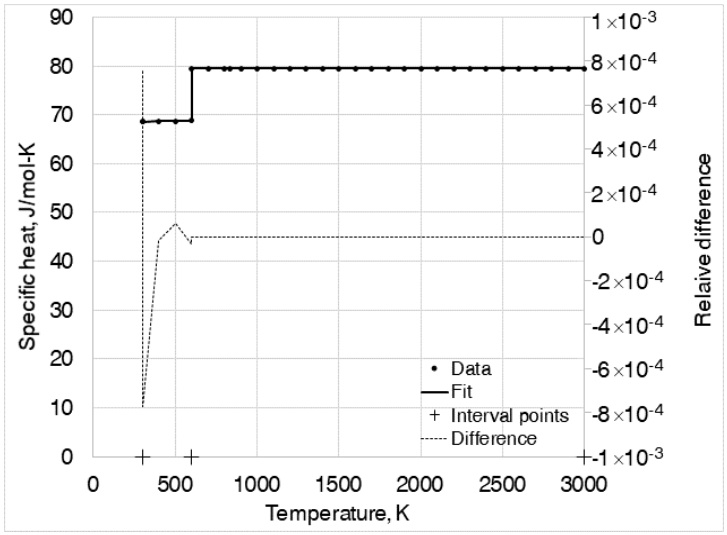


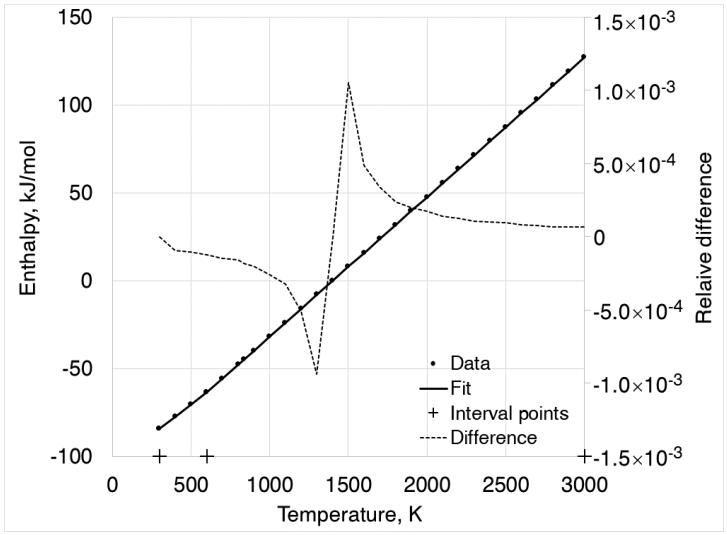


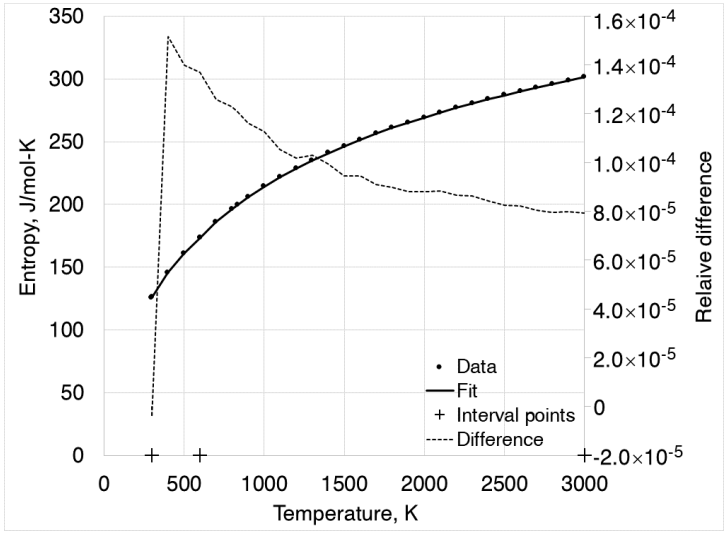


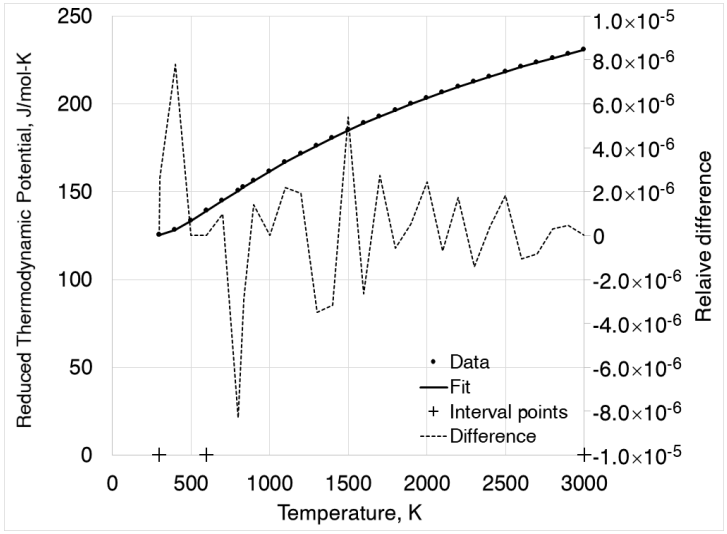
Sodium Cyanide (NaCN) – C1N1Na1(l)

MW = 49.007 g/mol – $\Delta_f H^o$(298.15 K) = -84.398 kJ/mol – S^o(298.15 K) = 125.507 J/mol-K – p^o = 0.1 MPa



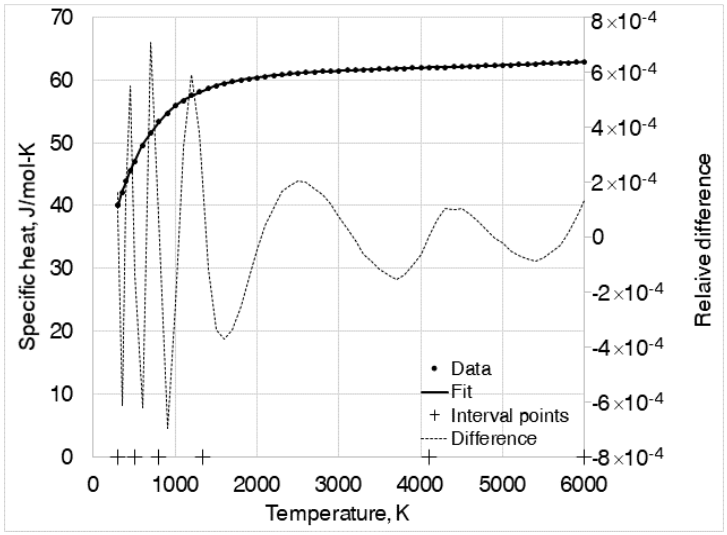


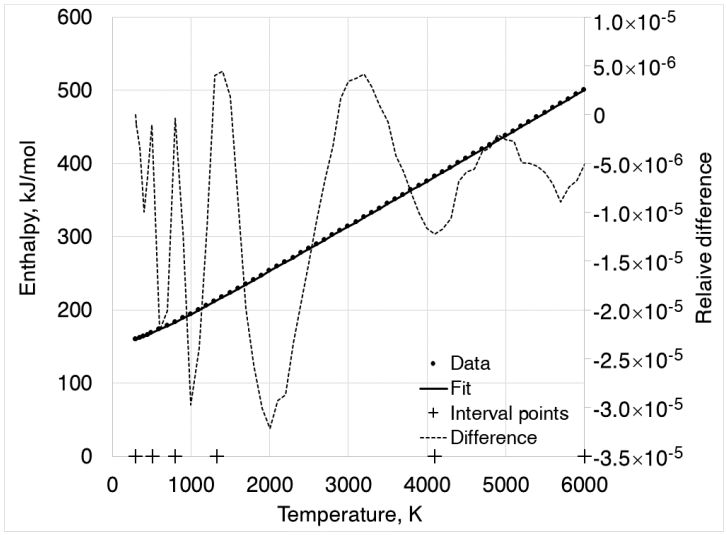


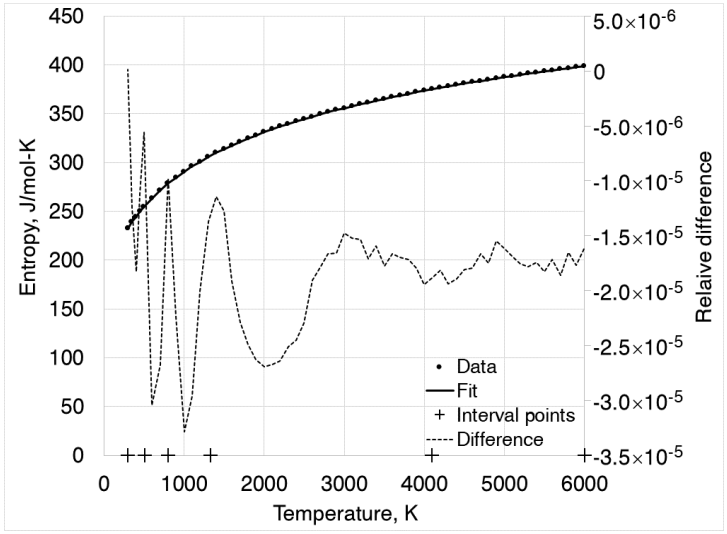


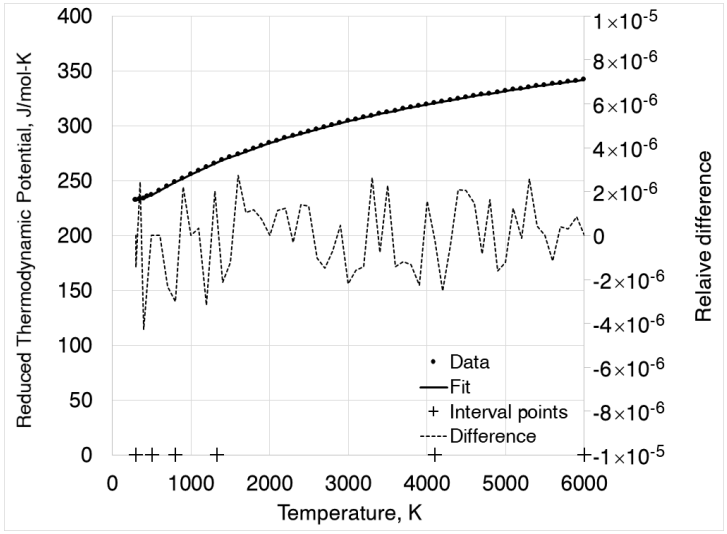
NCO Radical (NC_{101}) – C1N1O1(g)

MW = 42.017 g/mol – $\Delta_f H^o$(298.15 K) = 159.41 kJ/mol – S^o(298.15 K) = 232.248 J/mol-K – p^o = 0.1 MPa



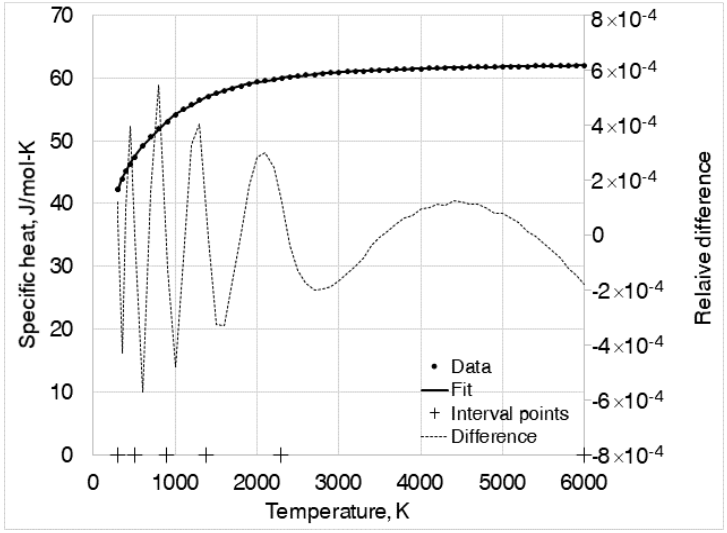


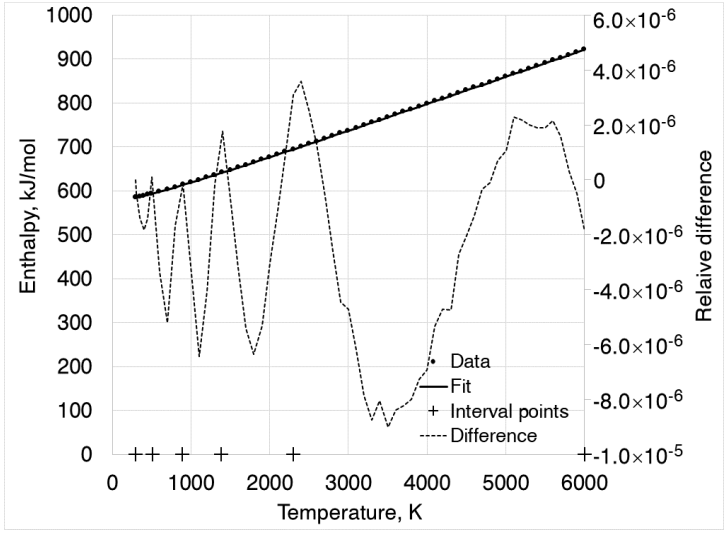


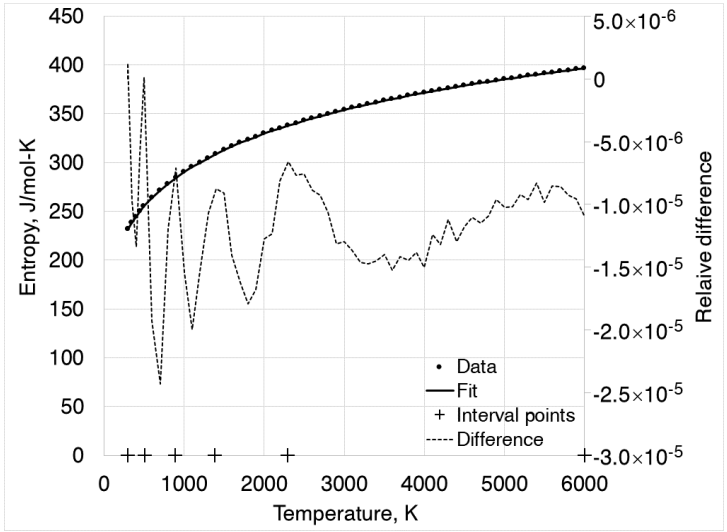


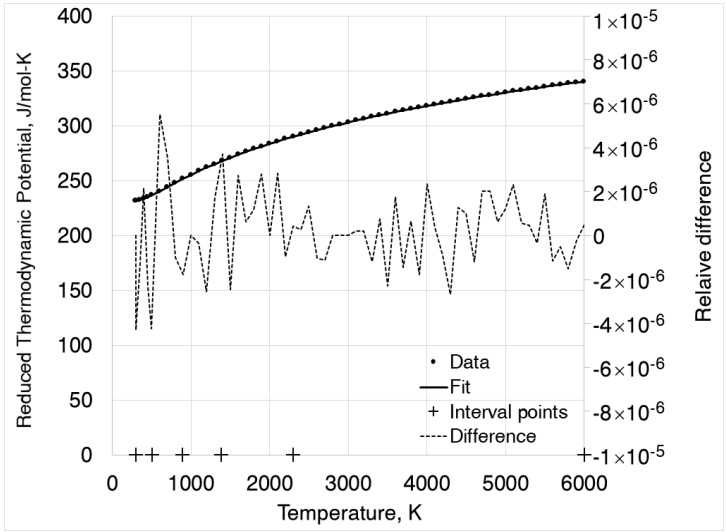
CNN Radical (CNN) – C1N2(g)_01

MW = 40.024 g/mol – $\Delta_f H^o$(298.15 K) = 584.505 kJ/mol – S^o(298.15 K) = 231.716 J/mol-K – p^o = 0.1 MPa



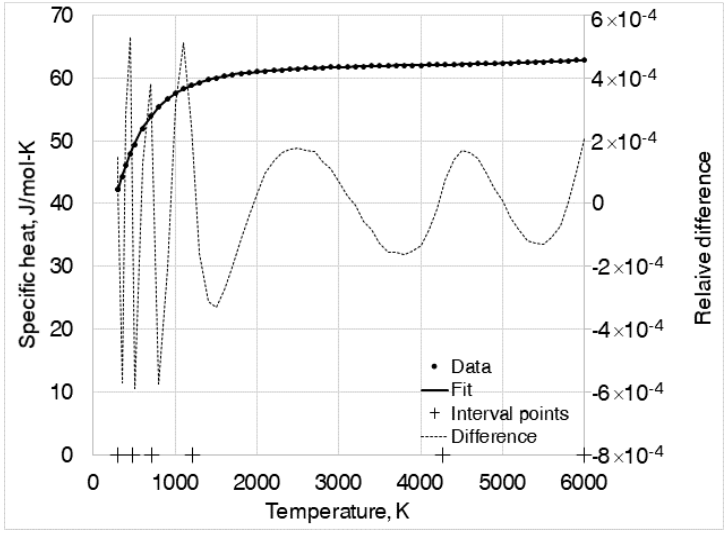


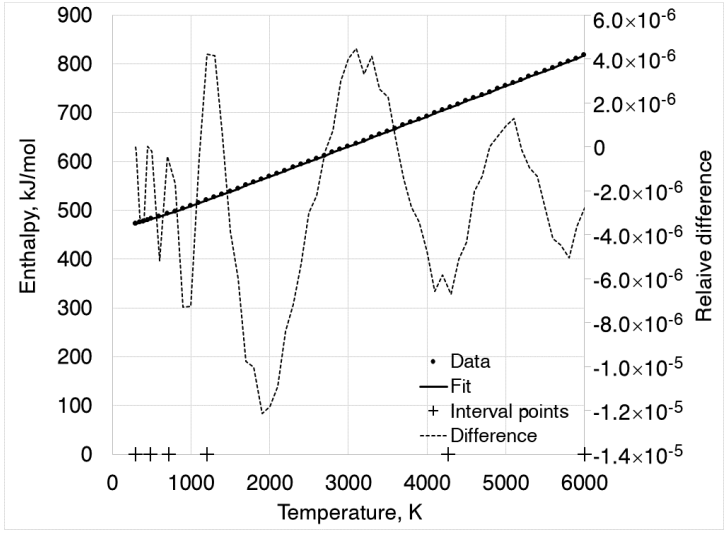


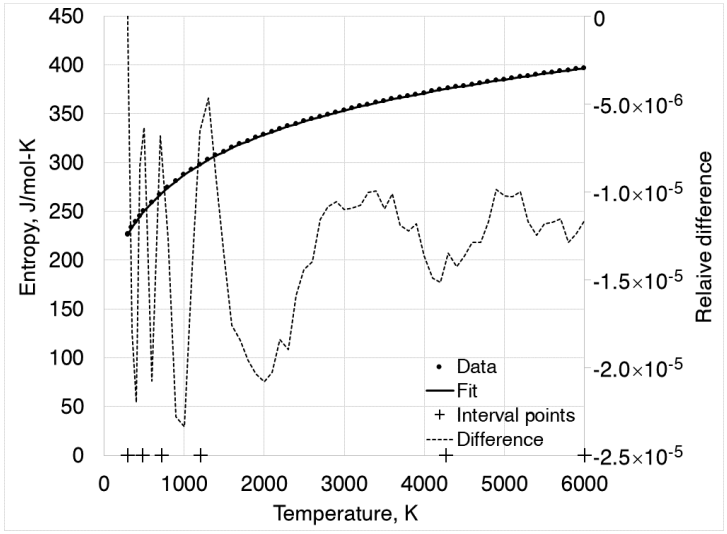


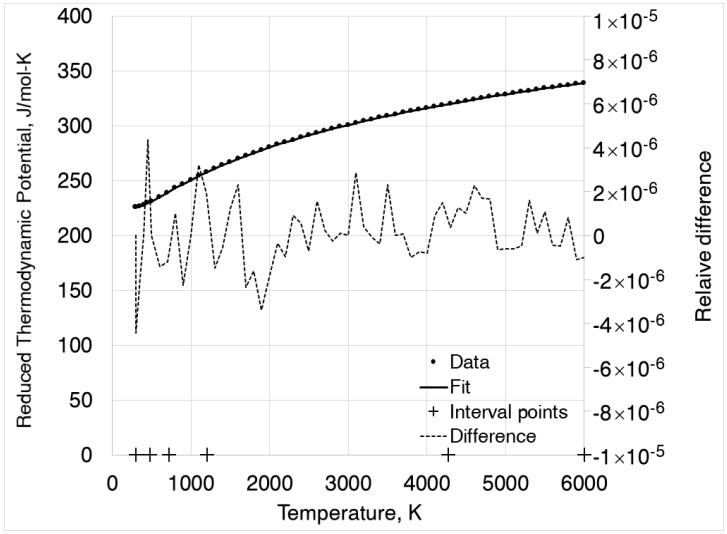
NCN Radical (NCN) – C1N2(g)_02

MW = 40.024 g/mol – $\Delta_f H^o$(298.15 K) = 472.792 kJ/mol – S^o(298.15 K) = 226.203 J/mol-K – p^o = 0.1 MPa



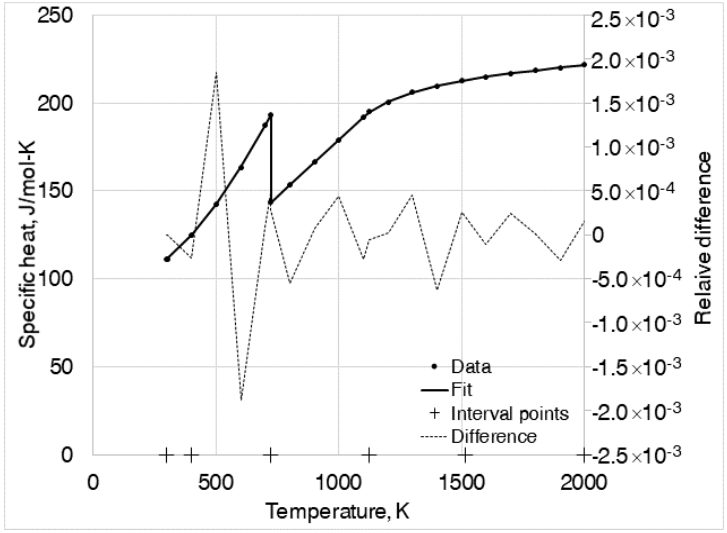


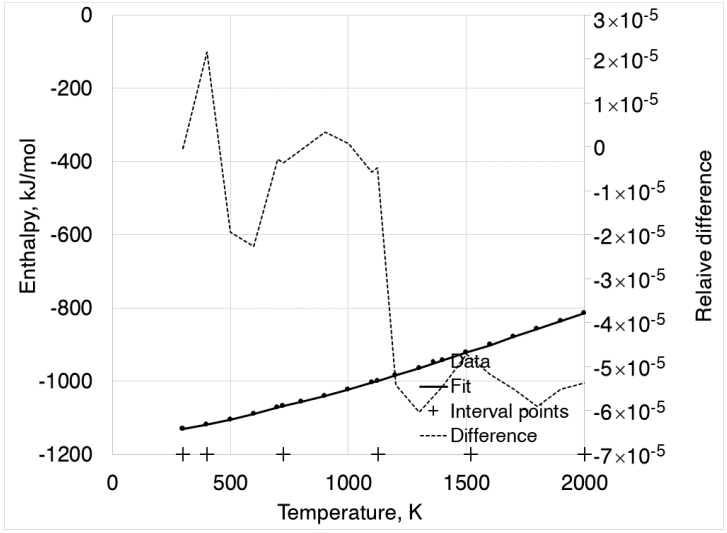


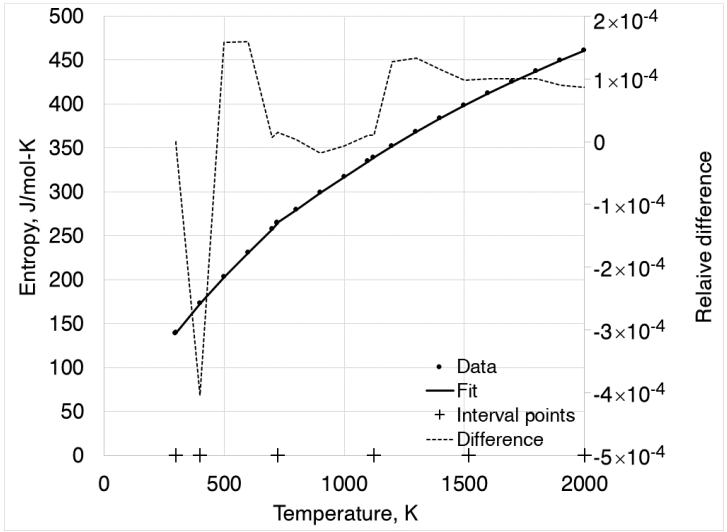


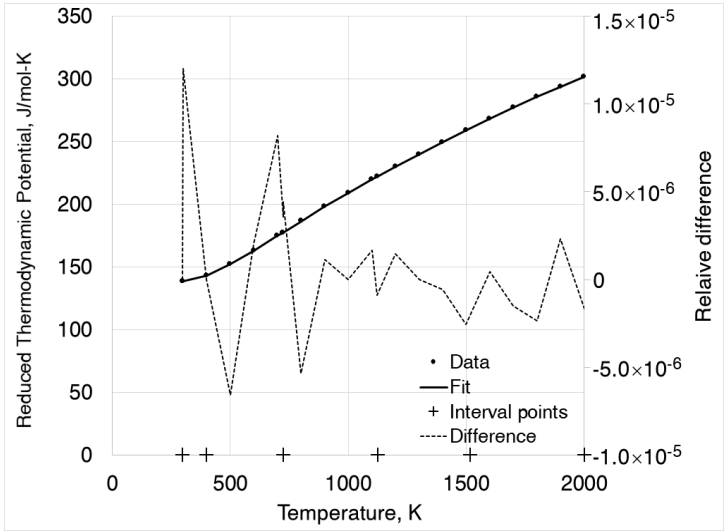
Sodium Carbonate (Na_2CO_3) – C1Na2O3(cr)

MW = 105.989 g/mol – $\Delta_f H^o$(298.15 K) = -1130.768 kJ/mol – S^o(298.15 K) = 138.797 J/mol-K – p^o = 0.1 MPa



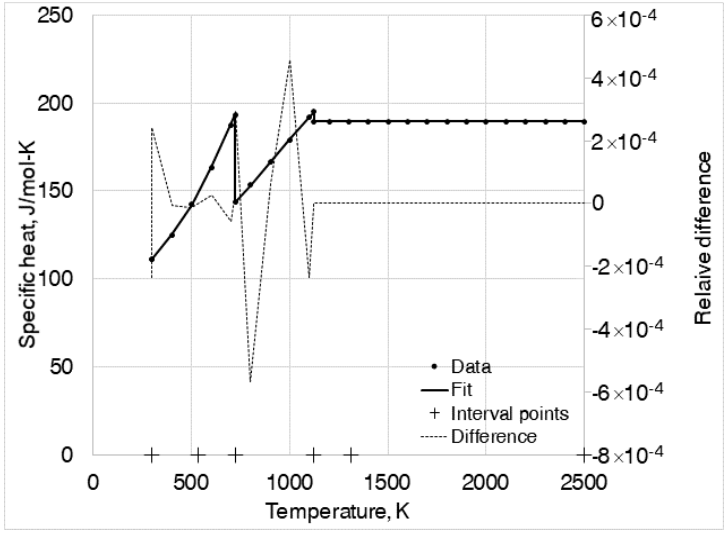


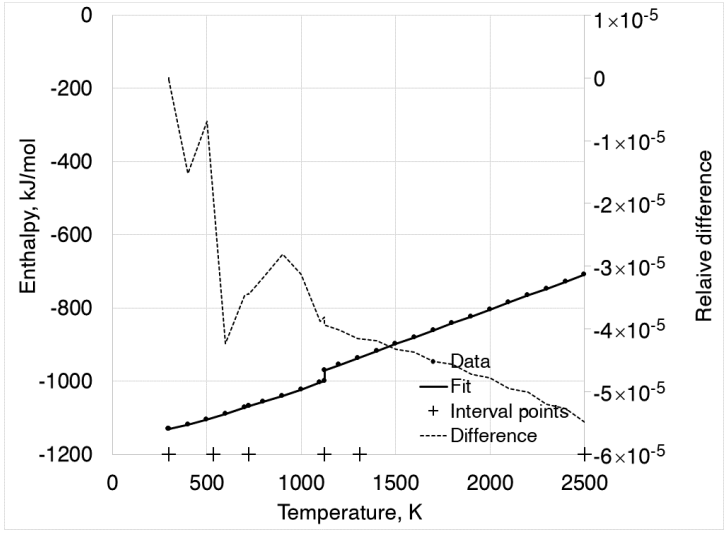


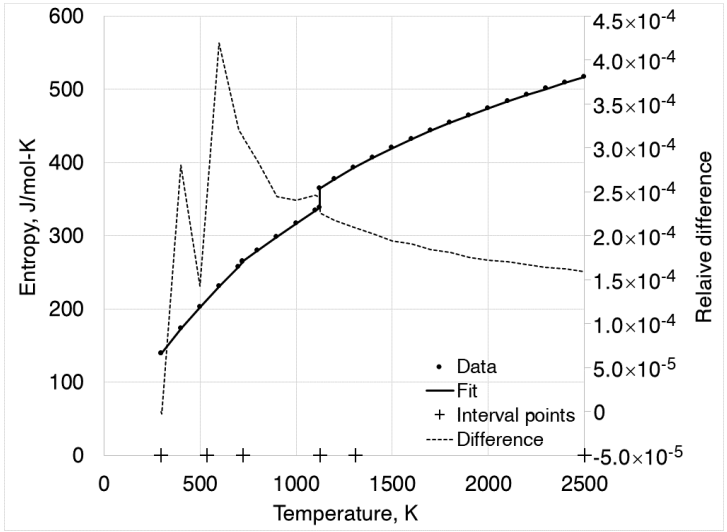


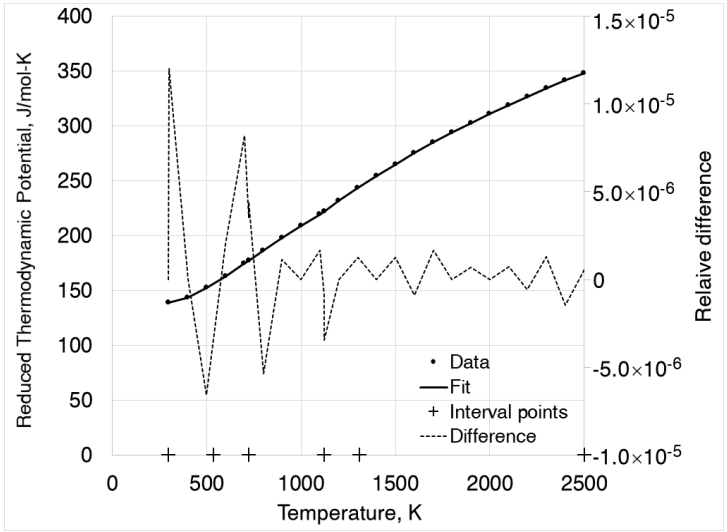
Sodium Carbonate ($NaCO_3$) – C1Na2O3(cr,l)

MW = 105.989 g/mol – $\Delta_f H^o$(298.15 K) = -1130.768 kJ/mol – S^o(298.15 K) = 138.797 J/mol-K – p^o = 0.1 MPa



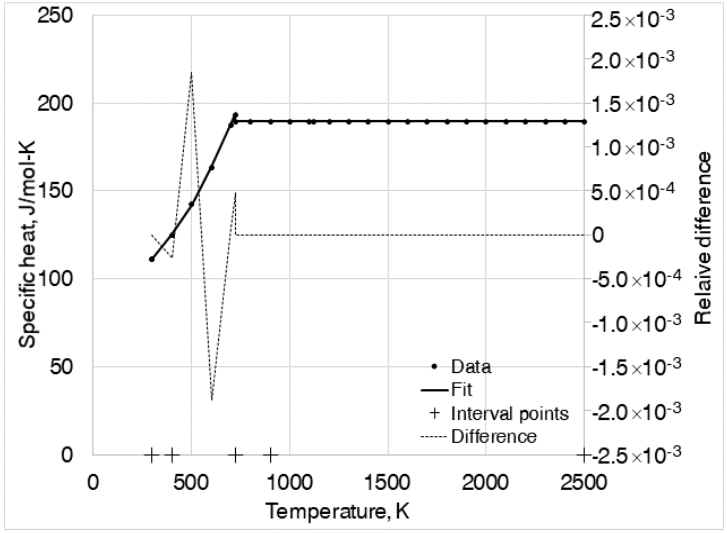


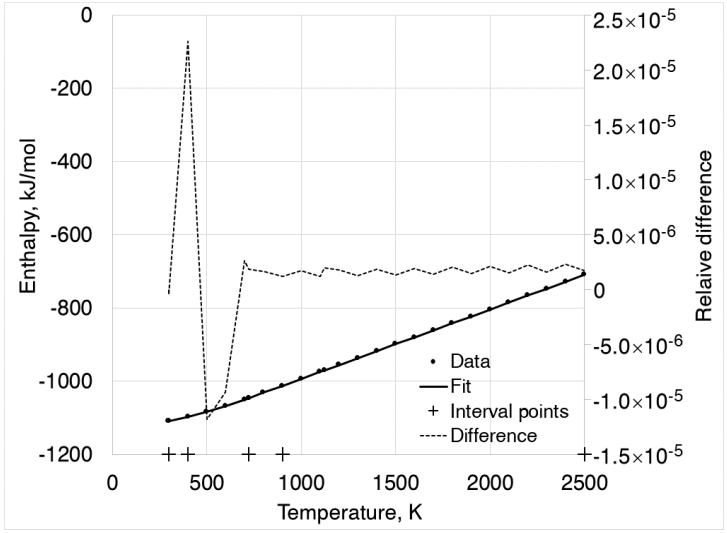


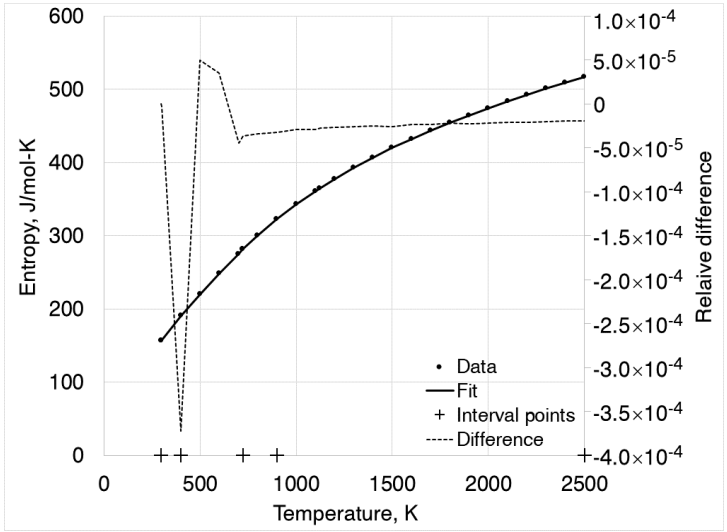


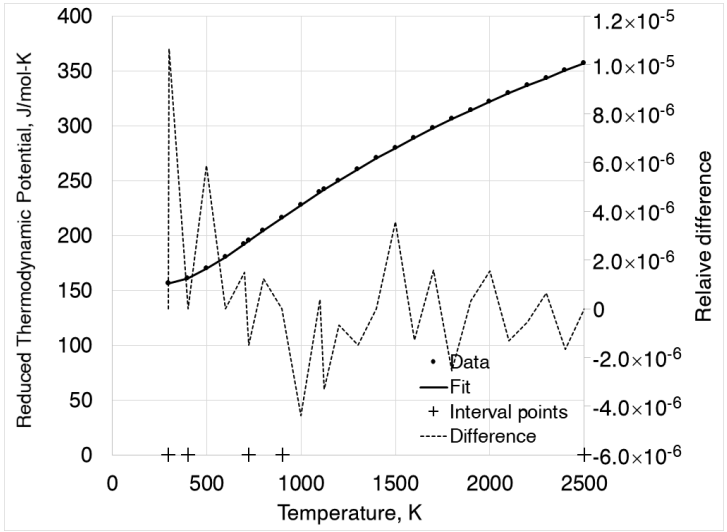
Sodium Carbonate (Na_2CO_3) – C1Na2O3(l)

MW = 105.989 g/mol – $\Delta_f H^o$(298.15 K) = -1108.515 kJ/mol – S^o(298.15 K) = 156.423 J/mol-K – p^o = 0.1 MPa



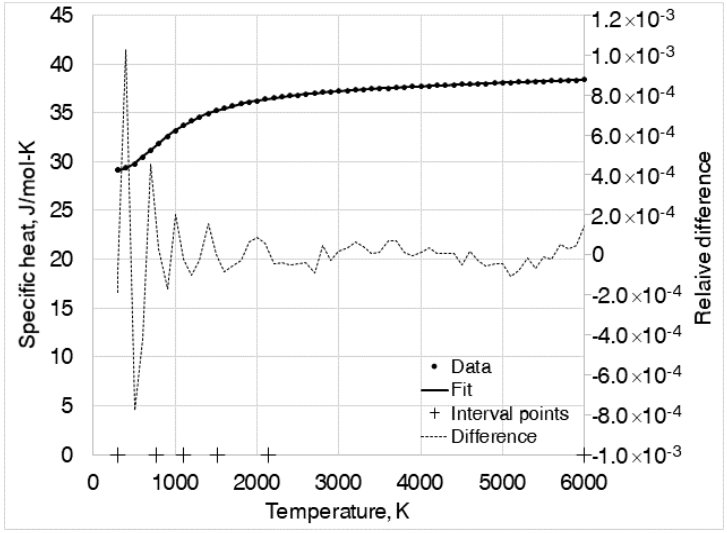


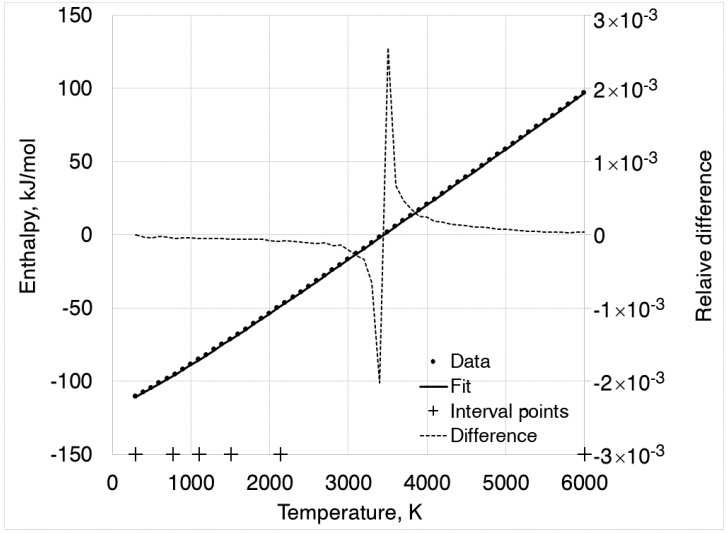


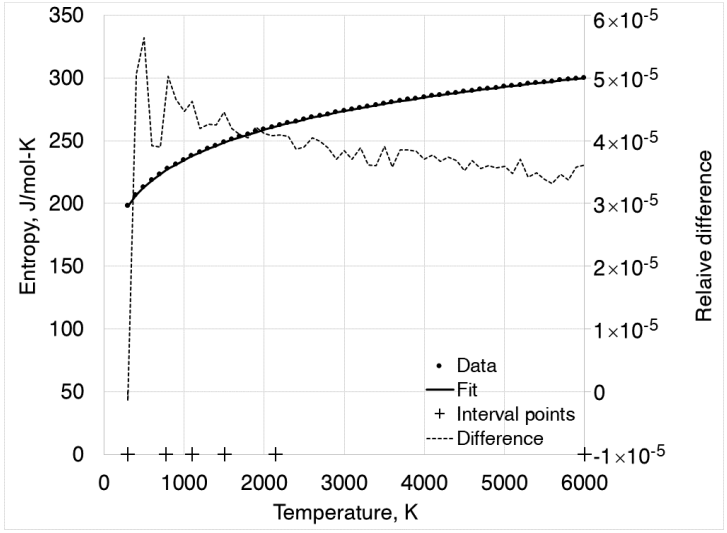


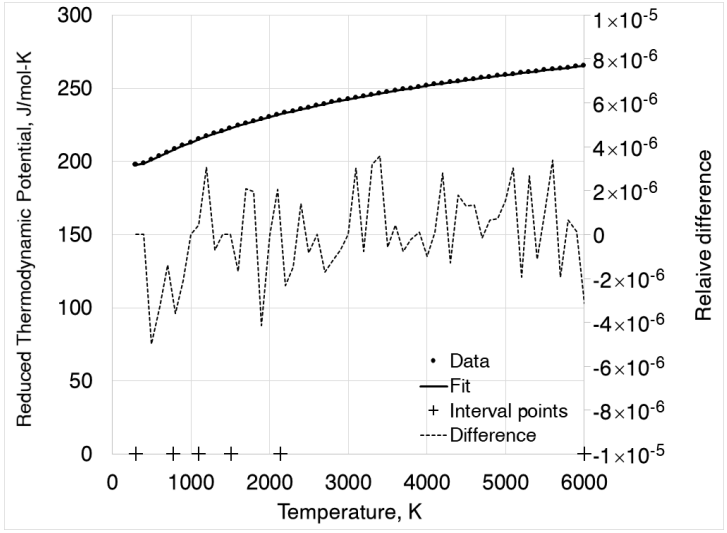
Carbon Monoxide (CO) – C1O1(g)

MW = 28.01 g/mol – $\Delta_f H^o$(298.15 K) = -110.527 kJ/mol – S^o(298.15 K) = 197.653 J/mol-K – p^o = 0.1 MPa



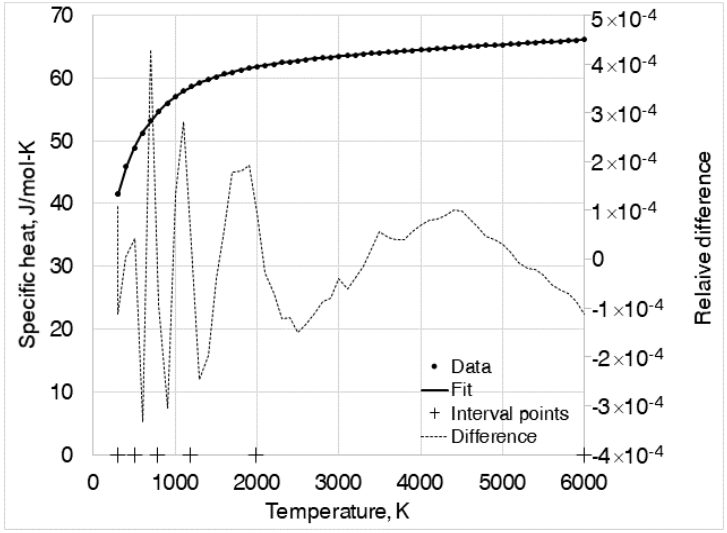


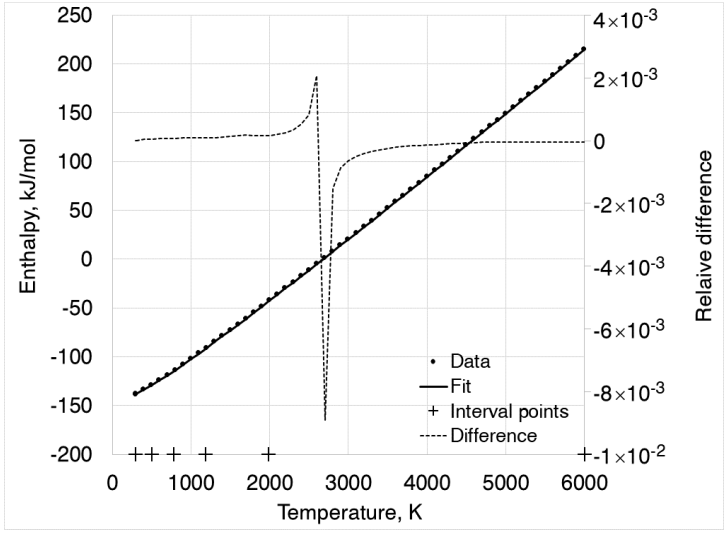


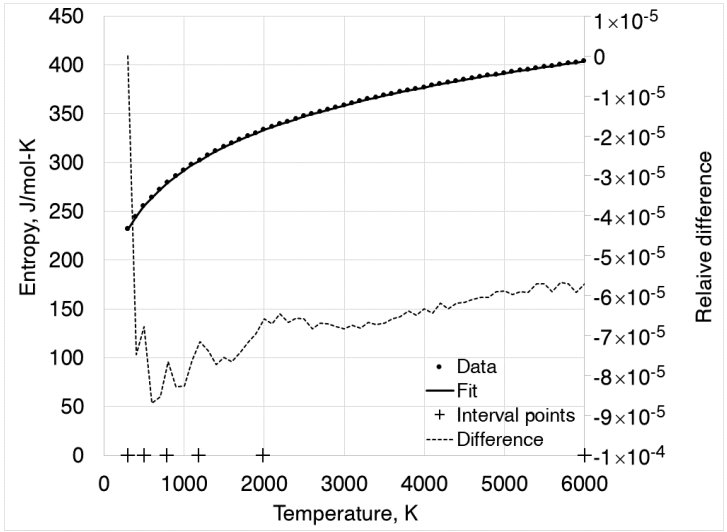


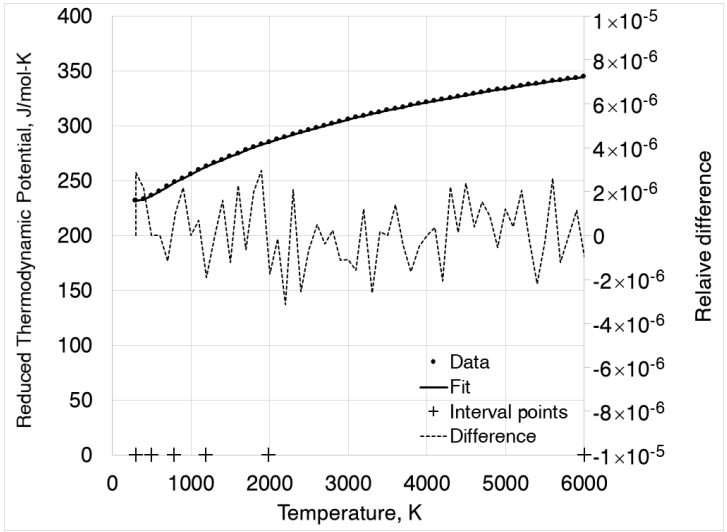
Carbon Oxide Sulfide (COS) – C1O1S1(g)

MW = 60.07 g/mol – $\Delta_f H^o$(298.15 K) = -138.407 kJ/mol – S^o(298.15 K) = 231.581 J/mol-K – p^o = 0.1 MPa



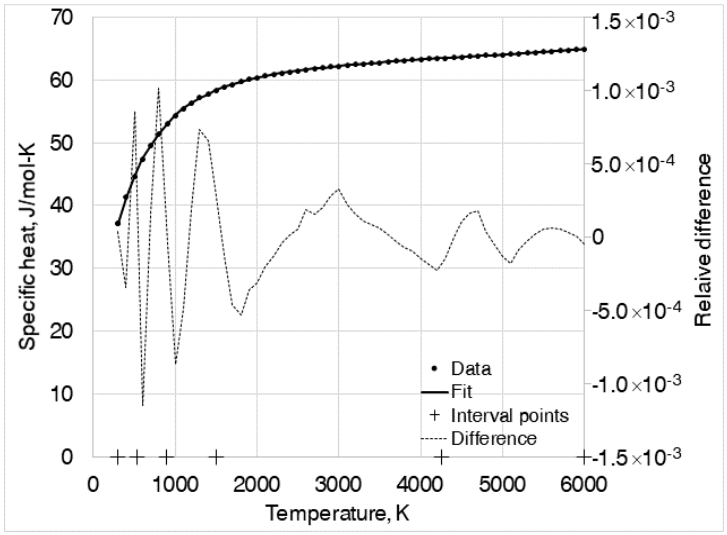


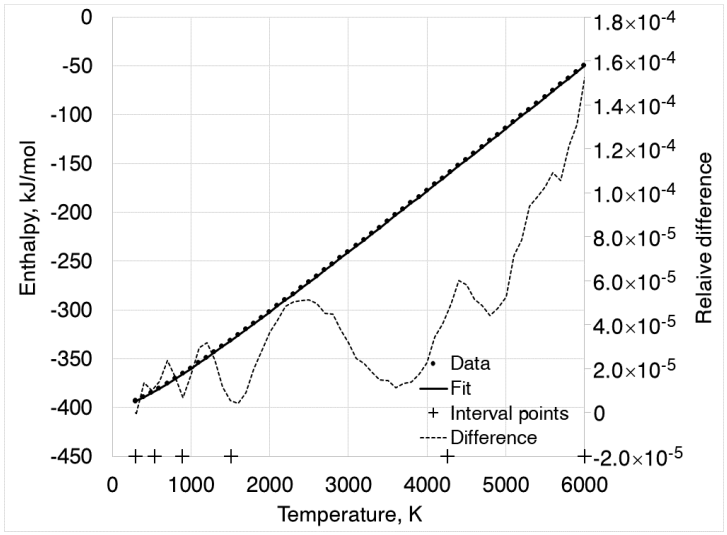


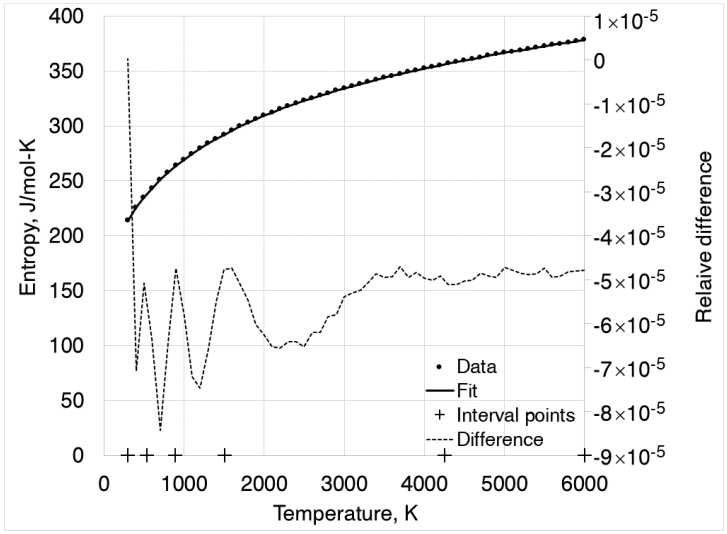


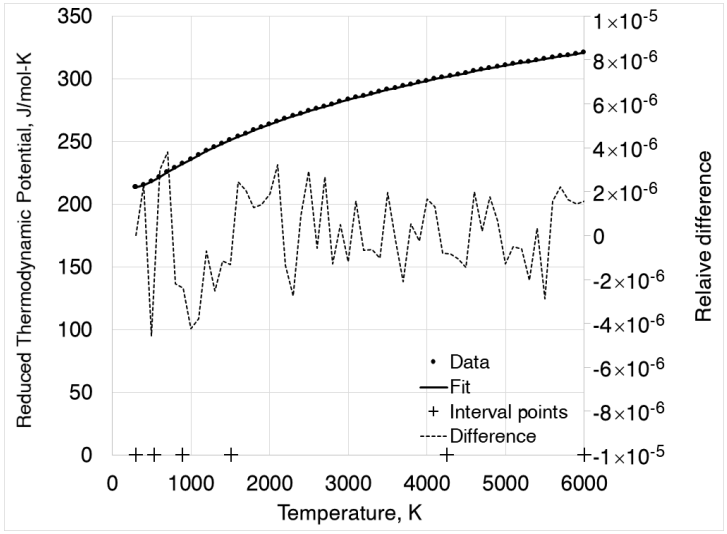
Carbon Dioxide (CO_2) – C1O2(g)

MW = 44.01 g/mol – $\Delta_f H^o$(298.15 K) = -393.522 kJ/mol – S^o(298.15 K) = 213.795 J/mol-K – p^o = 0.1 MPa



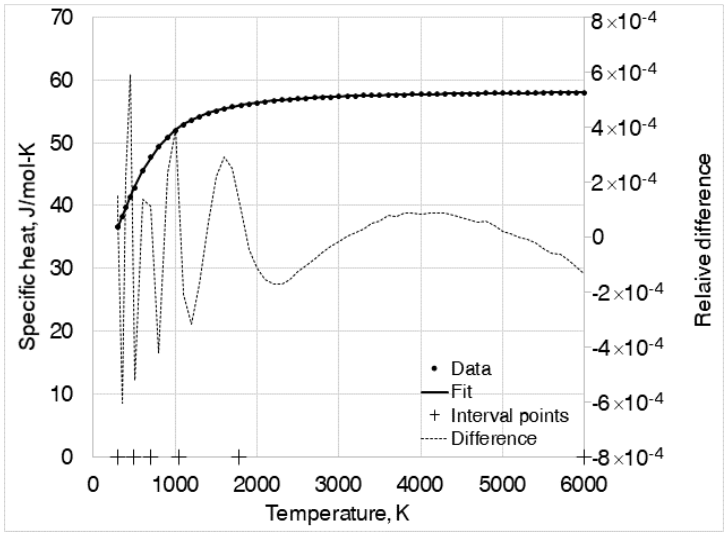


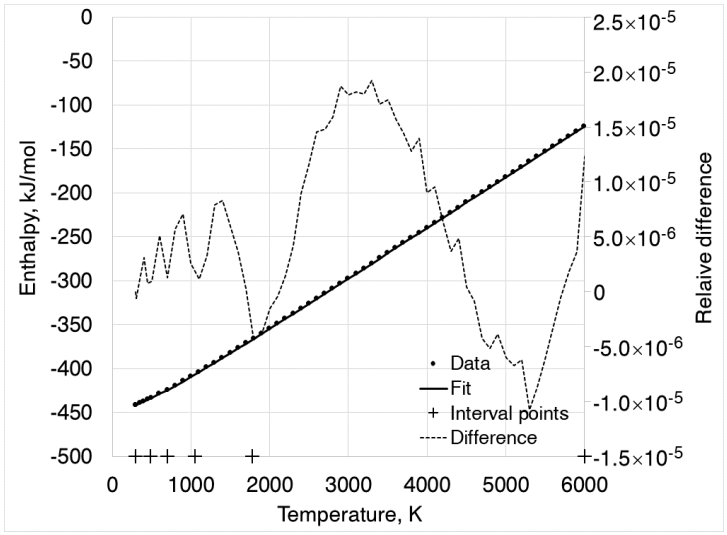


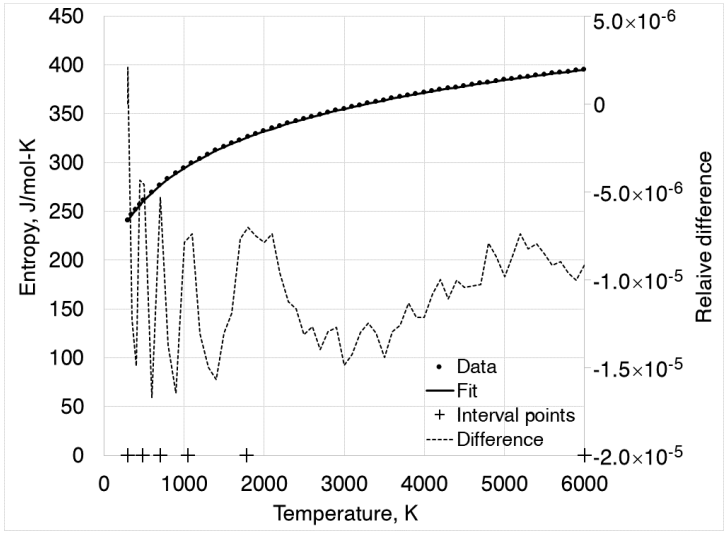


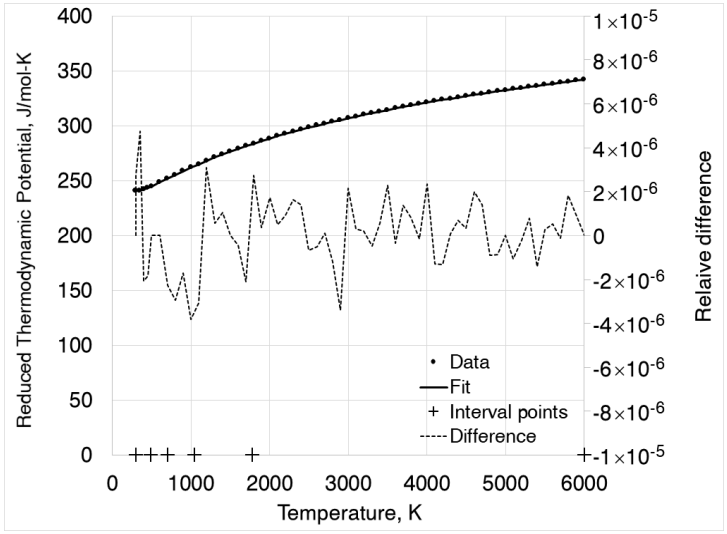
Carbon Dioxide, Ion (CO_2^-) – C1O2-(g)

MW = 44.01 g/mol – $\Delta_f H^o$(298.15 K) = -441.412 kJ/mol – S^o(298.15 K) = 240.648 J/mol-K – p^o = 0.1 MPa



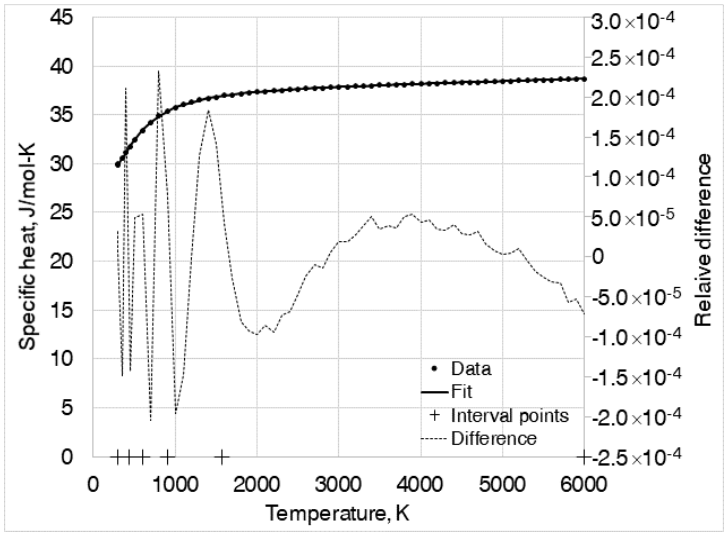


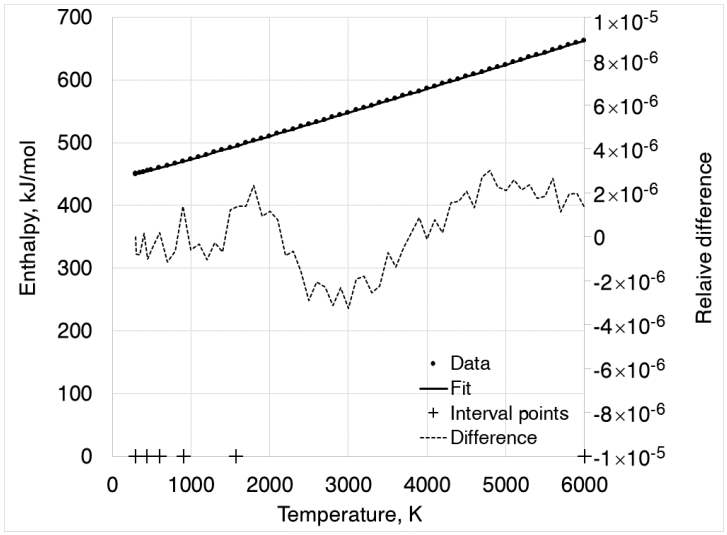


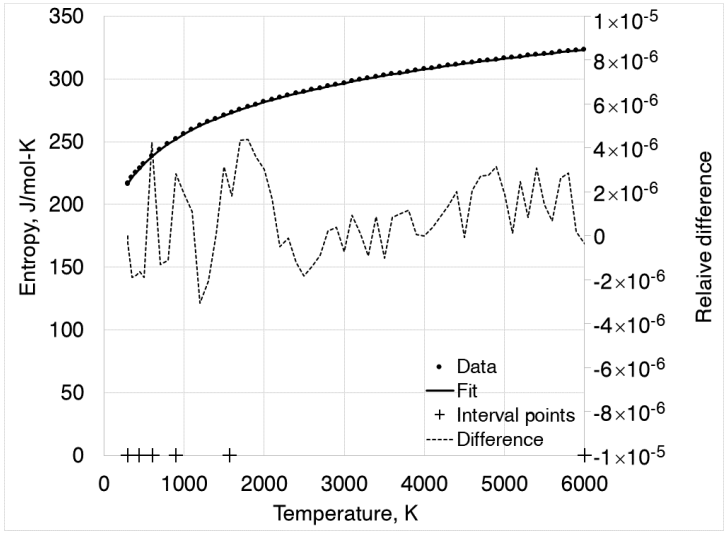


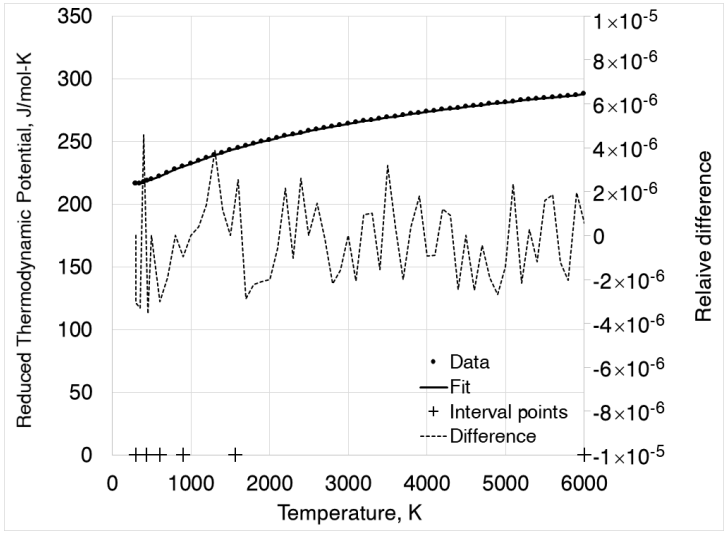
Carbon Phosphide (CP) – C1P1(g)

MW = 42.985 g/mol – $\Delta_f H^o$(298.15 K) = 449.893 kJ/mol – S^o(298.15 K) = 216.255 J/mol-K – p^o = 0.1 MPa



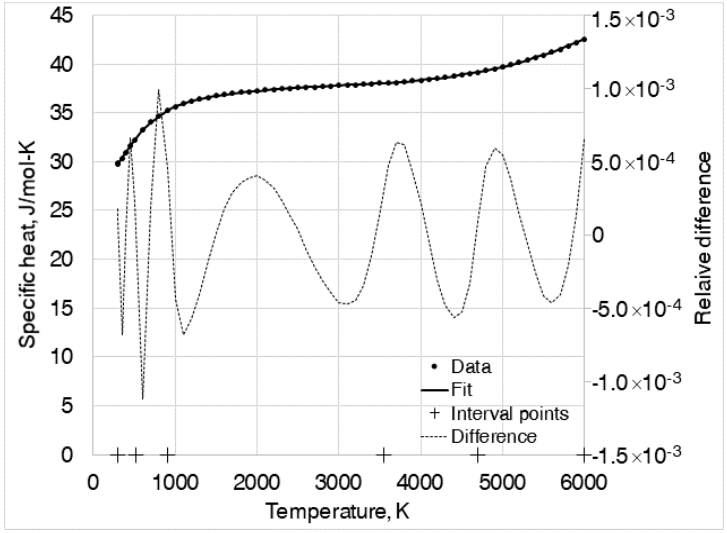


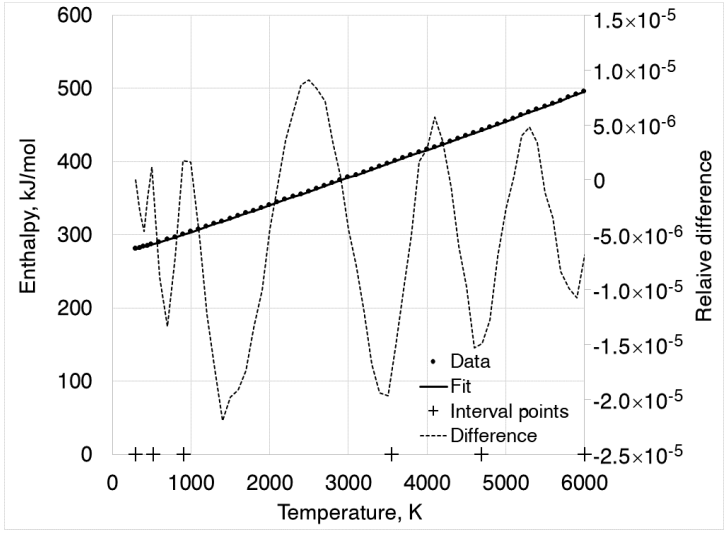


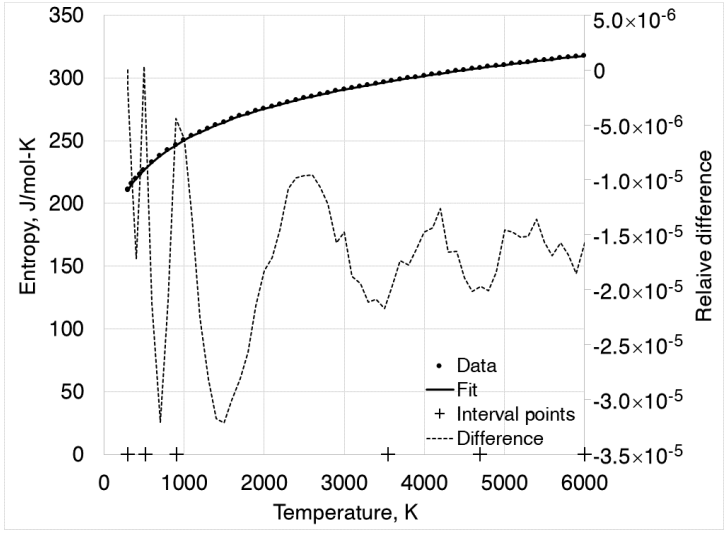


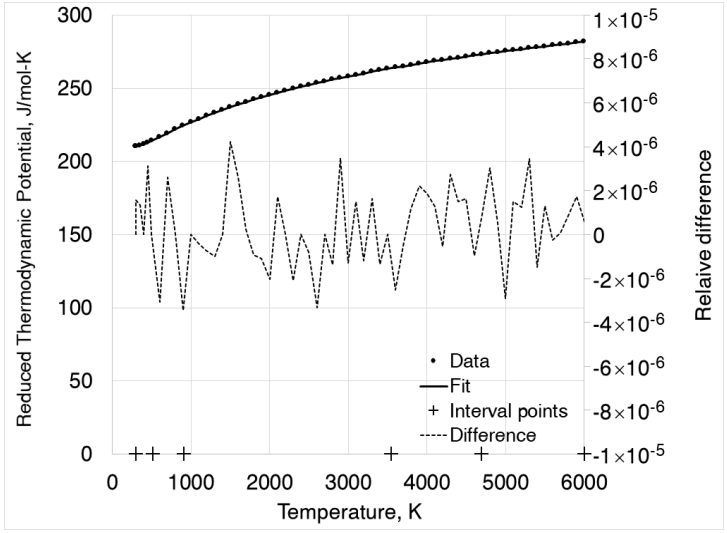
Carbon Sulfide (CS) – C1S1(g)

MW = 44.07 g/mol – $\Delta_f H^o$(298.15 K) = 280.328 kJ/mol – S^o(298.15 K) = 210.554 J/mol-K – p^o = 0.1 MPa



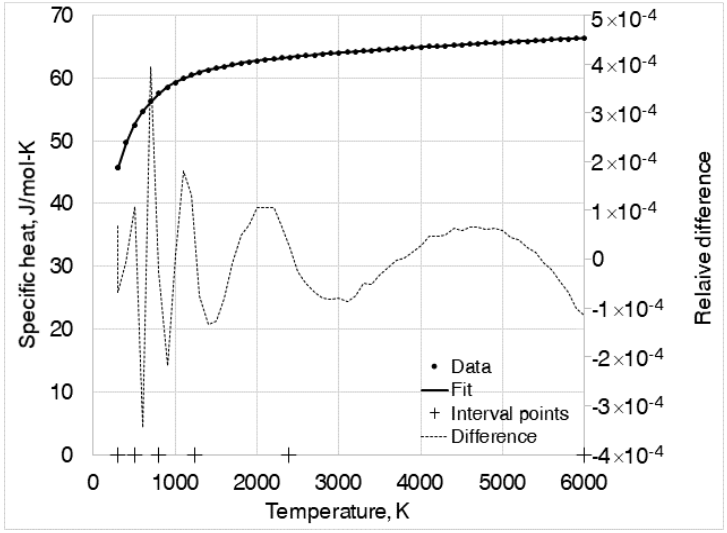


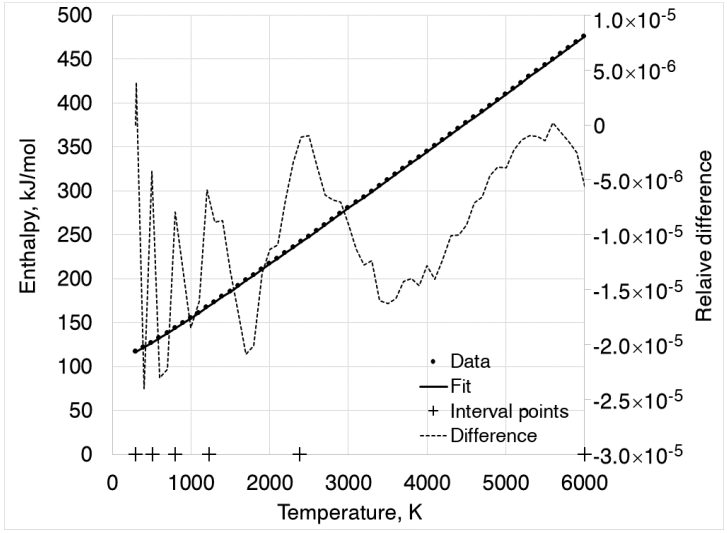


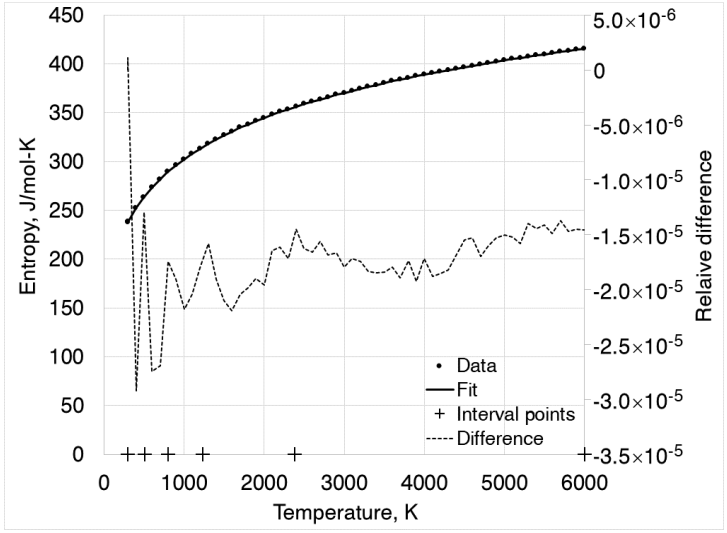


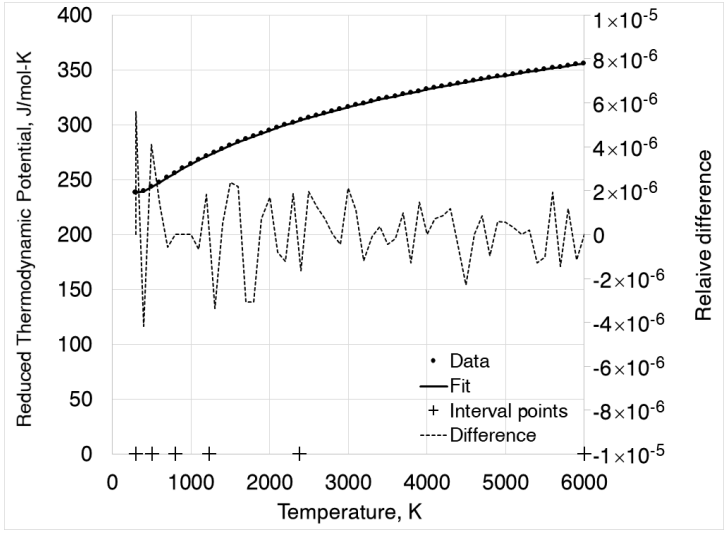
Carbon Disulfide (CS_2) – C1S2(g)

MW = 76.13 g/mol – $\Delta_f H^o$(298.15 K) = 116.943 kJ/mol – S^o(298.15 K) = 237.977 J/mol-K – p^o = 0.1 MPa



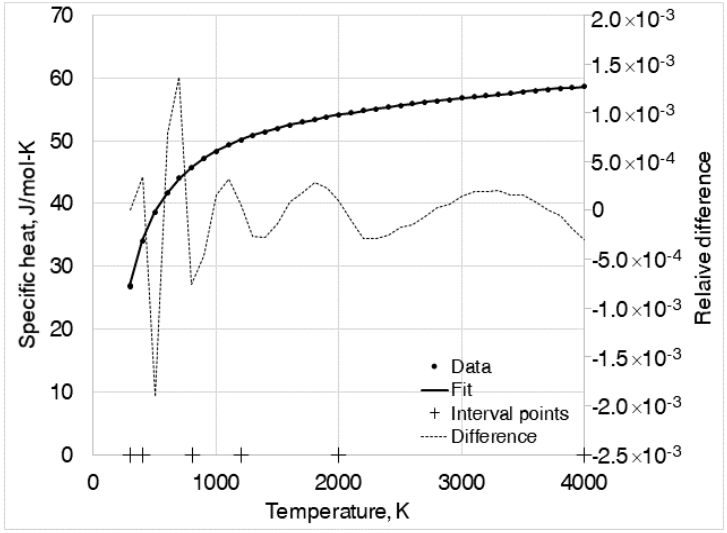


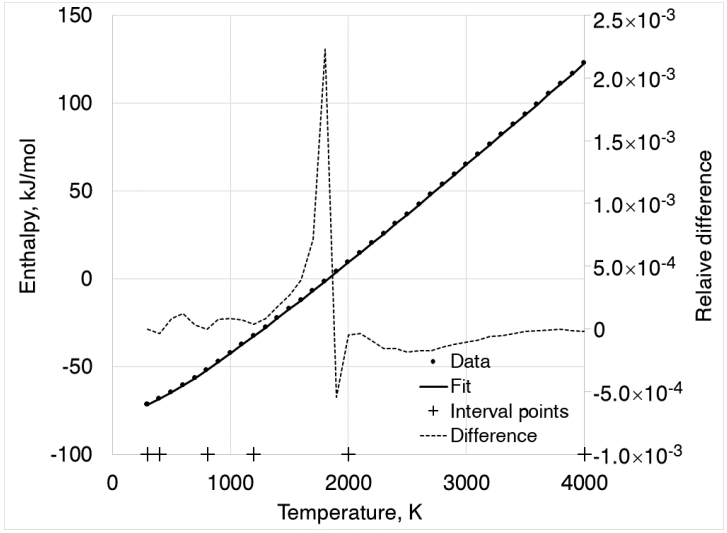


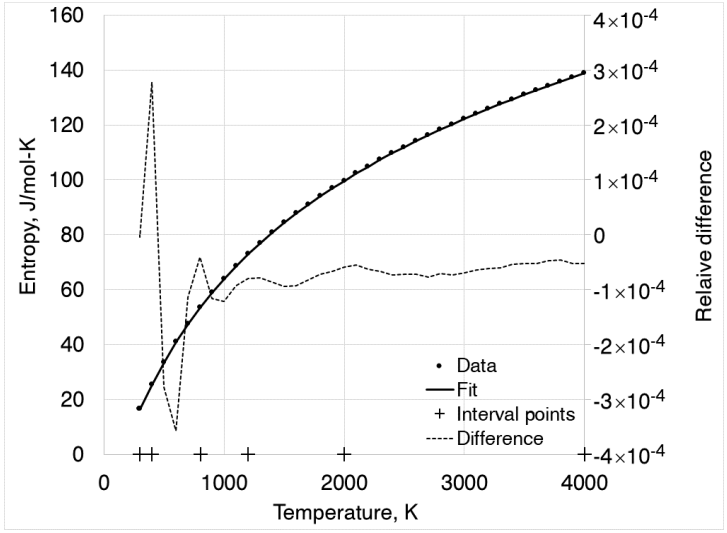


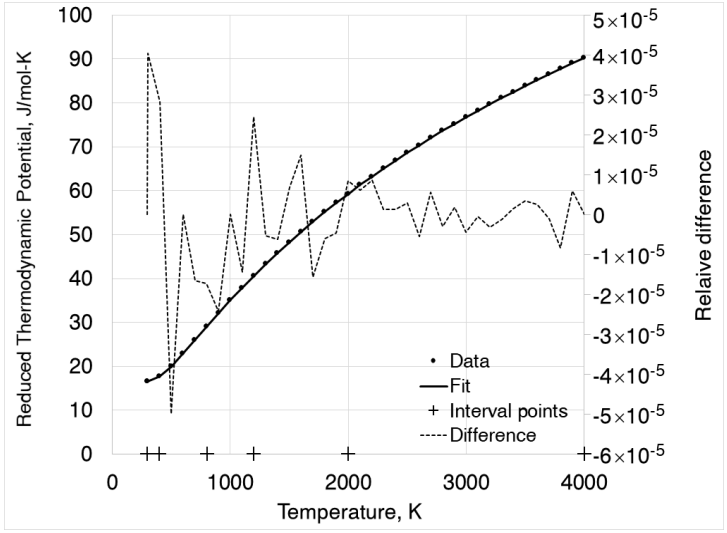
Silicon Carbide, Alpha (SiC) – C1Si1(cr)_01

MW = 40.096 g/mol – $\Delta_f H^o$(298.15 K) = -71.546 kJ/mol – S^o(298.15 K) = 16.485 J/mol-K – p^o = 0.1 MPa



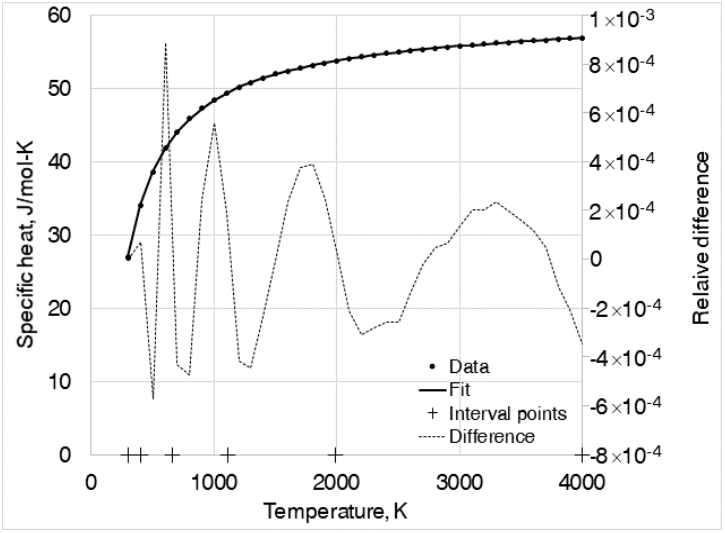


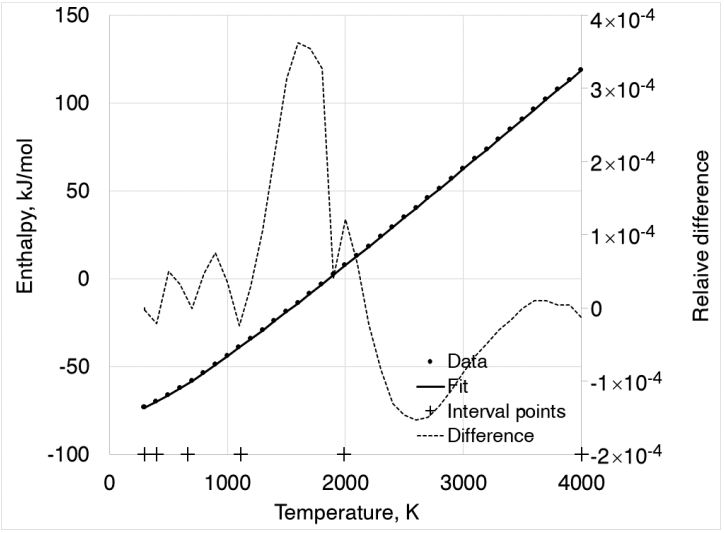


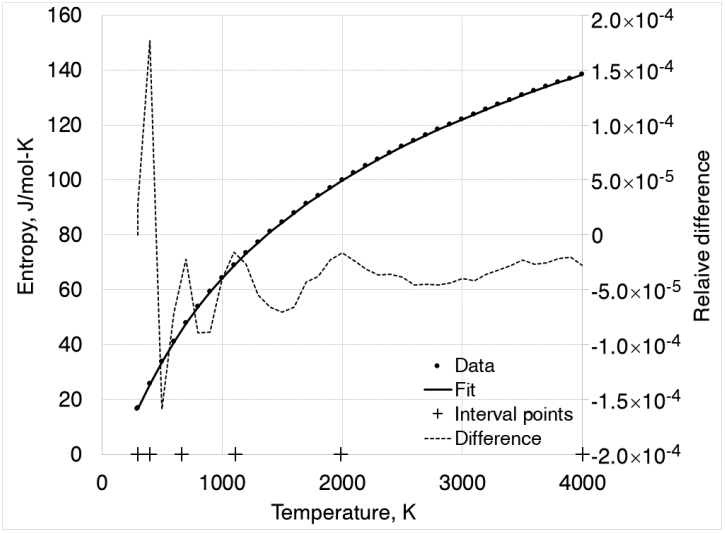


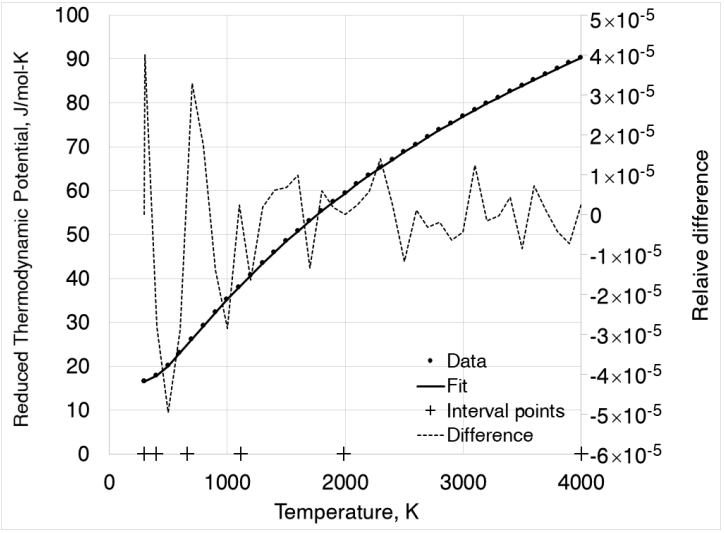
Silicon Carbide, Beta (SiC) – C1Si1(cr)_02

MW = 40.096 g/mol – $\Delta_f H^o$(298.15 K) = -73.22 kJ/mol – S^o(298.15 K) = 16.61 J/mol-K – p^o = 0.1 MPa



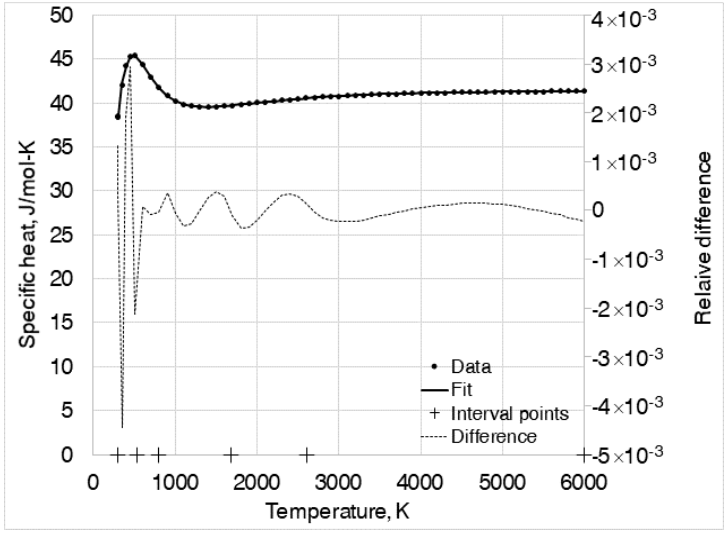


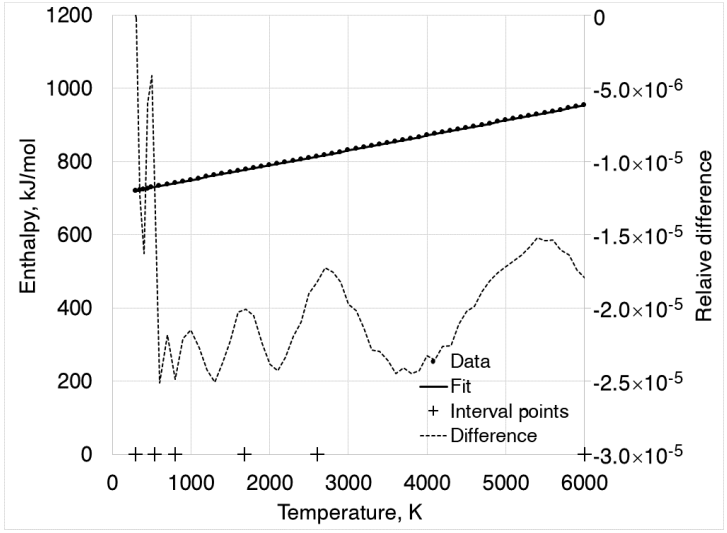


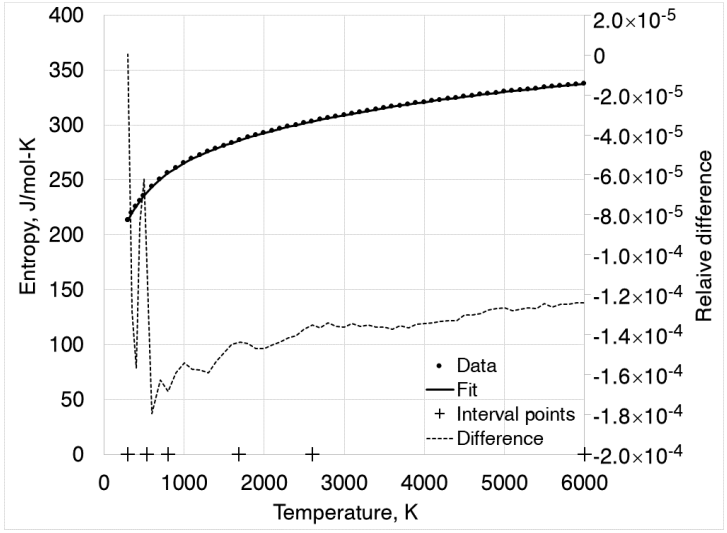


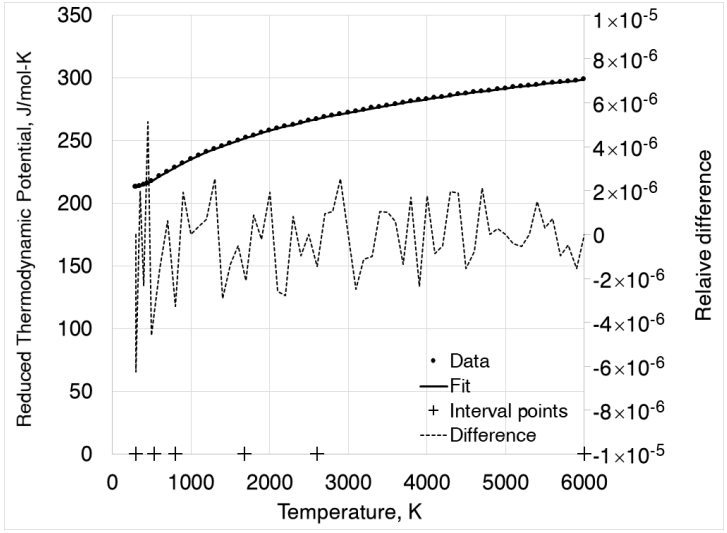
Silicon Carbide (SiC) – C1Si1(g)

MW = 40.096 g/mol – $\Delta_f H^o$(298.15 K) = 719.648 kJ/mol – S^o(298.15 K) = 213.044 J/mol-K – p^o = 0.1 MPa



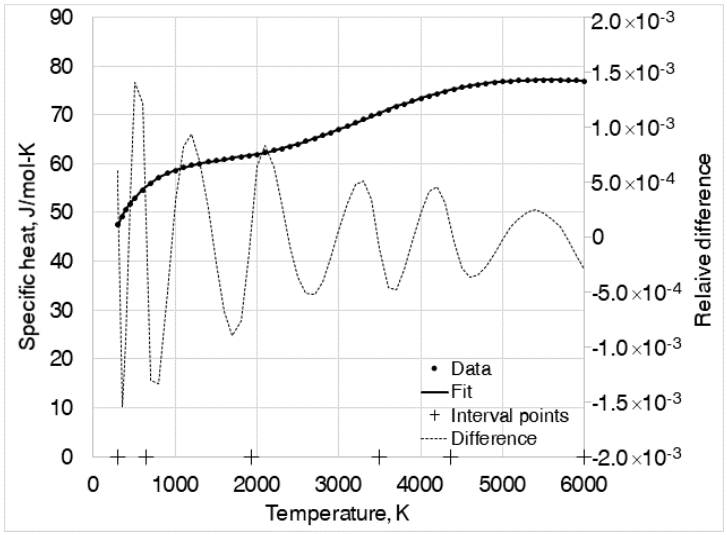


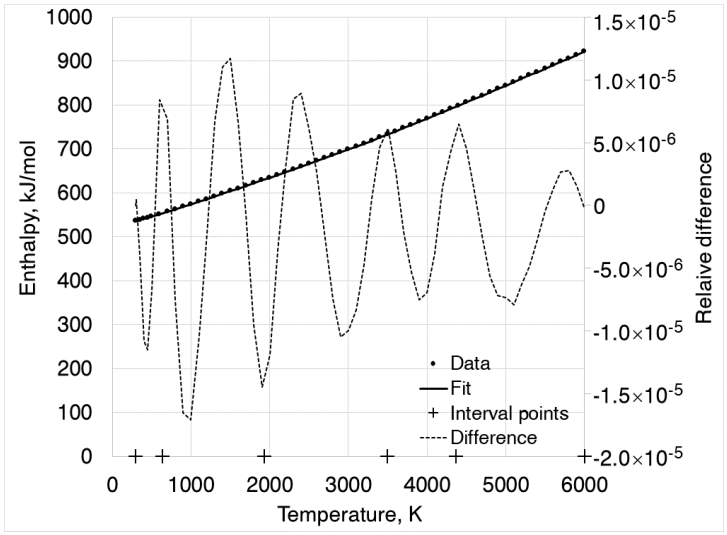


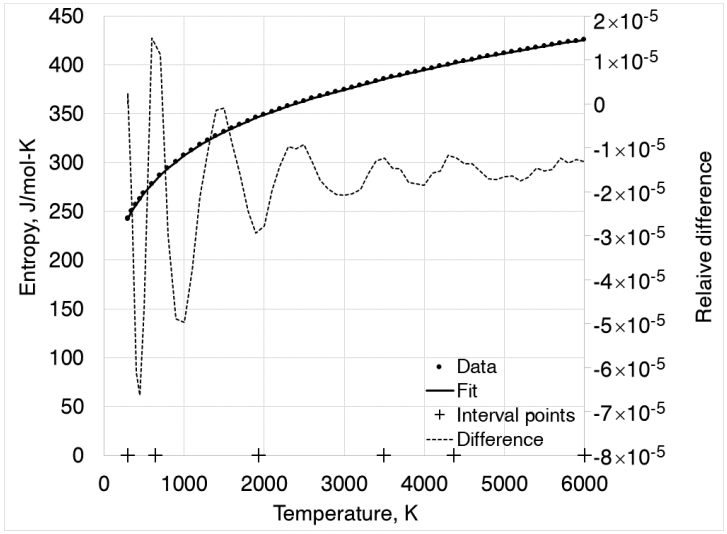


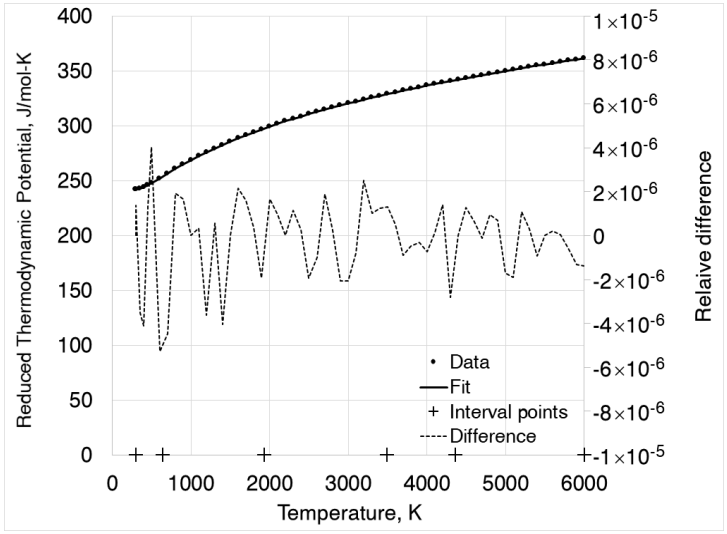
Silicon Carbide (Si_2C) – C1Si2(g)

MW = 68.182 g/mol – $\Delta_f H^o$(298.15 K) = 535.552 kJ/mol – S^o(298.15 K) = 242.286 J/mol-K – p^o = 0.1 MPa



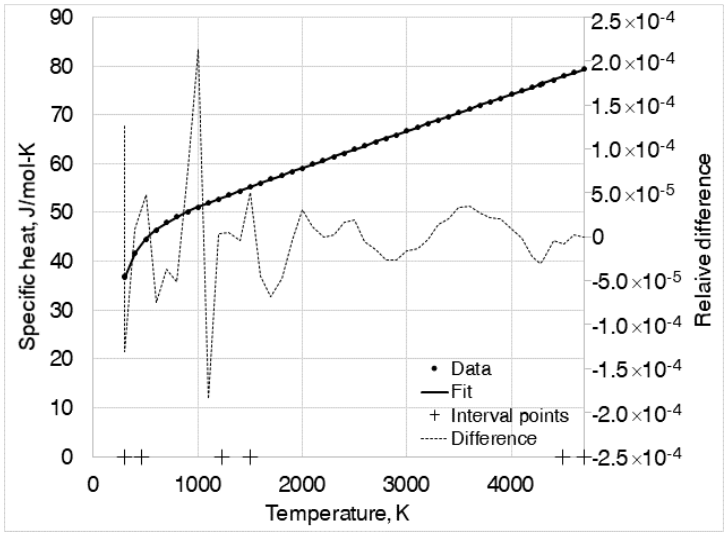


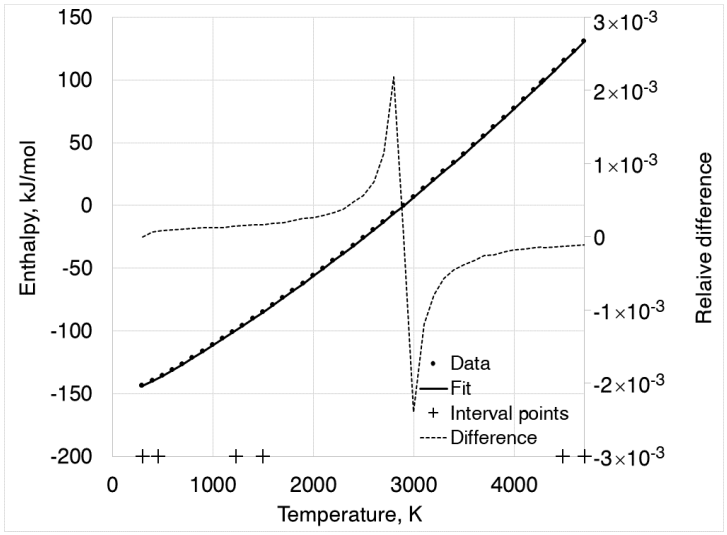


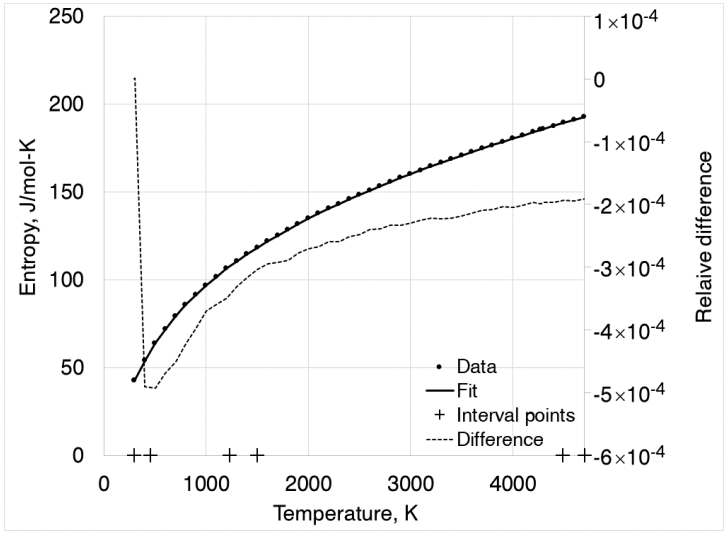


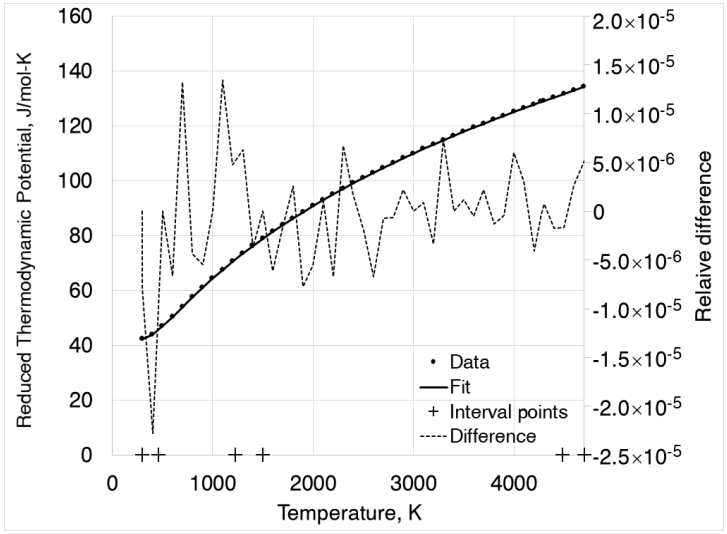
Tantalum Carbide (TaC) – C1Ta1(cr)

MW = 192.959 g/mol – $\Delta_f H^o$(298.15 K) = -144.097 kJ/mol – S^o(298.15 K) = 42.373 J/mol-K – p^o = 0.1 MPa



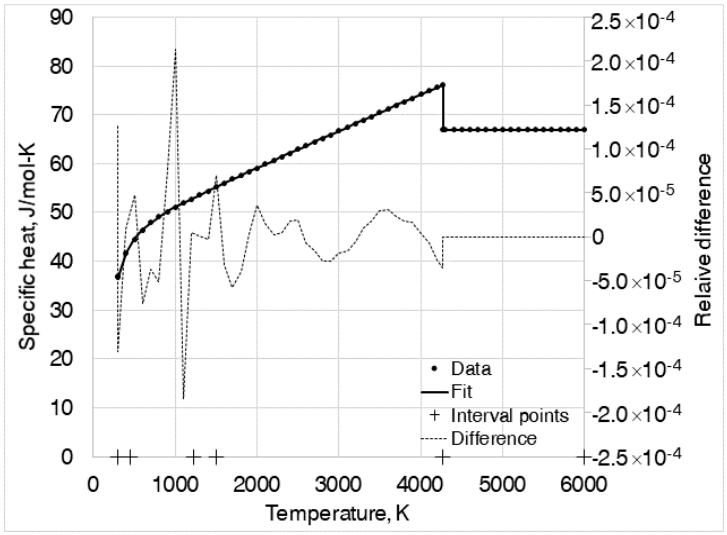


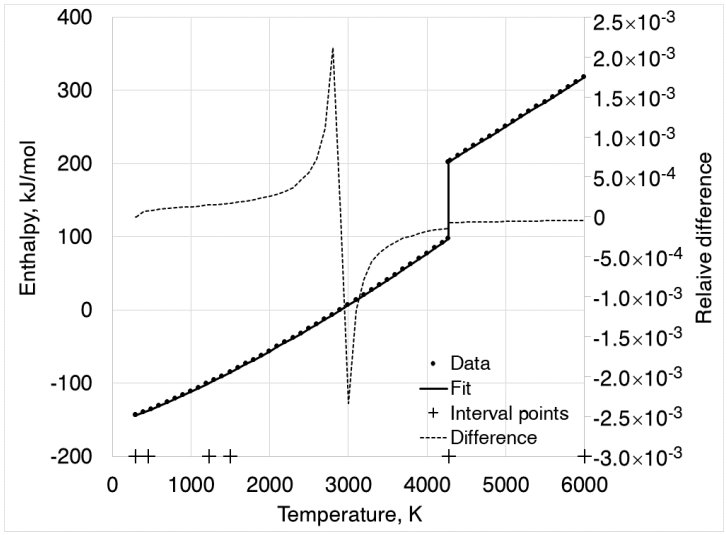


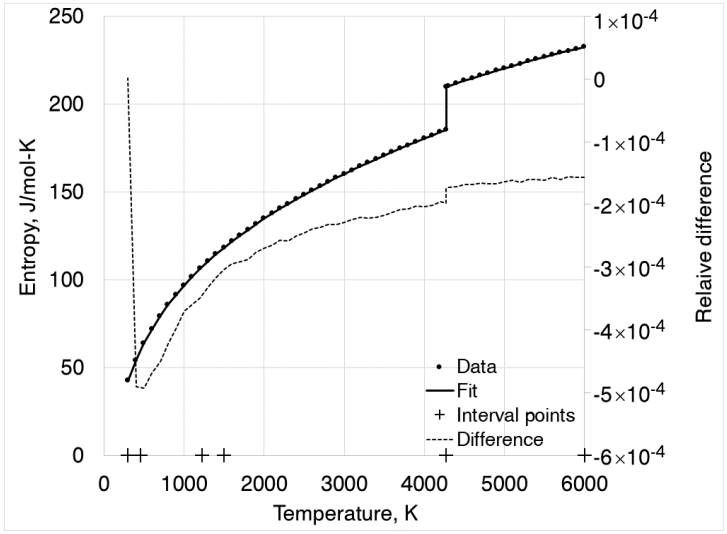


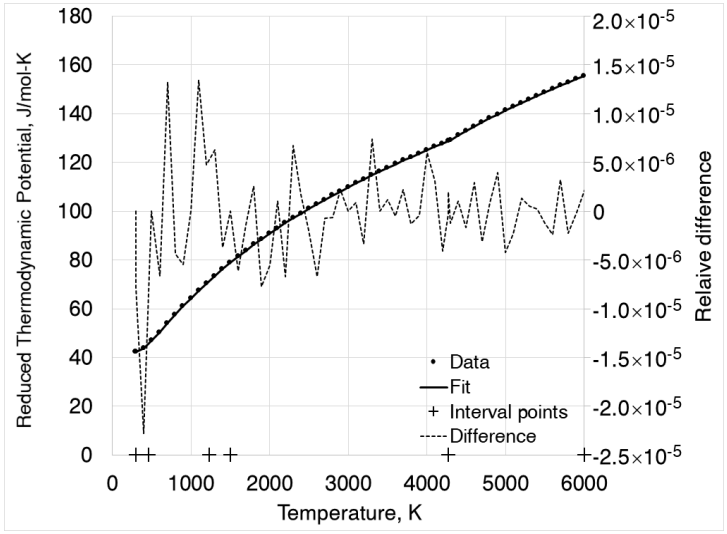
Tantalum Carbide (TaC) – C1Ta1(cr,l)

MW = 192.959 g/mol – $\Delta_f H^o$(298.15 K) = -144.097 kJ/mol – S^o(298.15 K) = 42.373 J/mol-K – p^o = 0.1 MPa



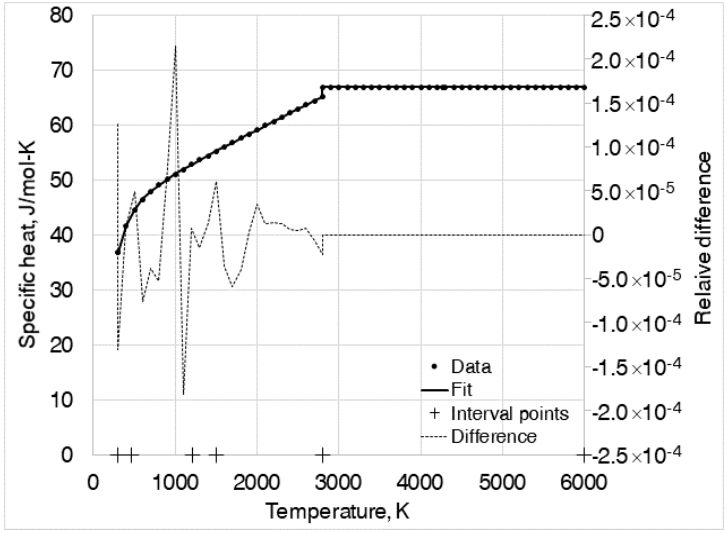


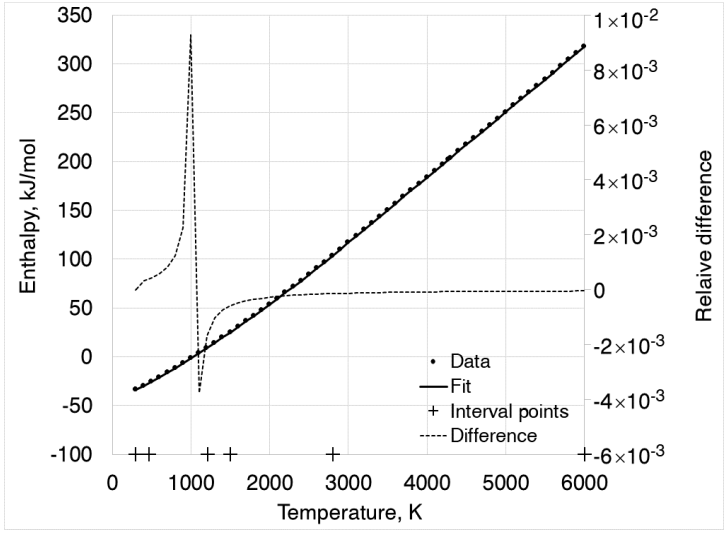


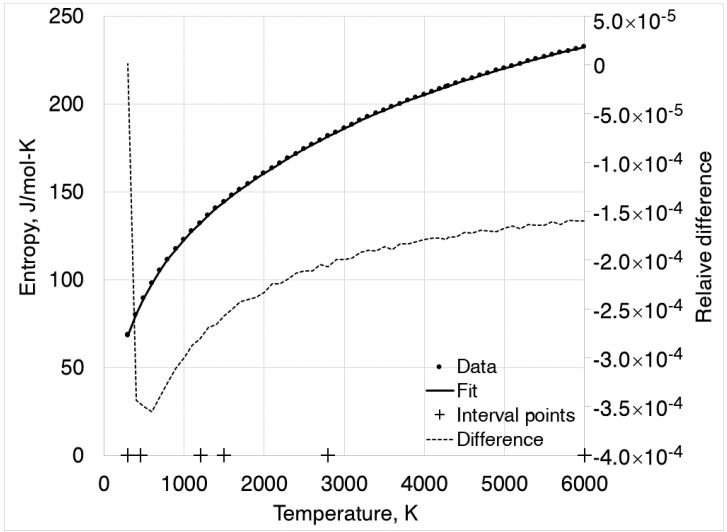


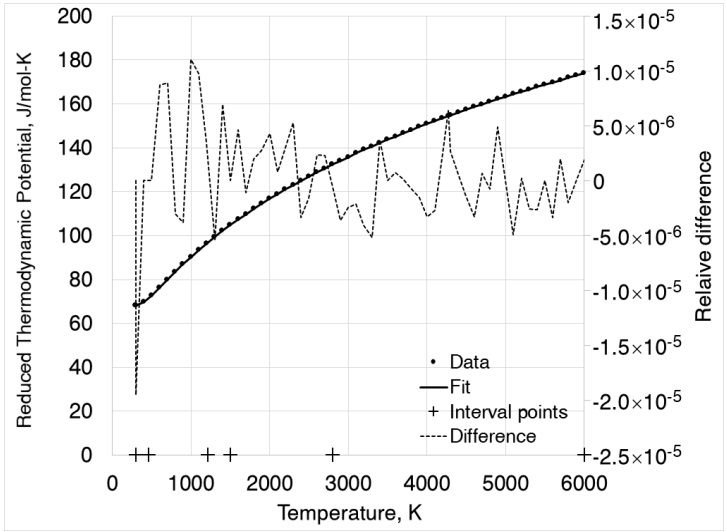
Tantalum Carbide (TaC) – C1Ta1(l)

MW = 192.959 g/mol – $\Delta_f H^o$(298.15 K) = -33.955 kJ/mol – S^o(298.15 K) = 68.279 J/mol-K – p^o = 0.1 MPa



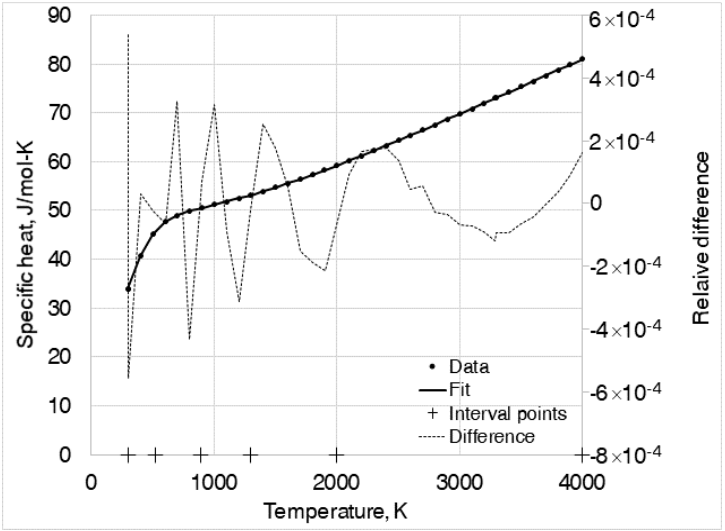


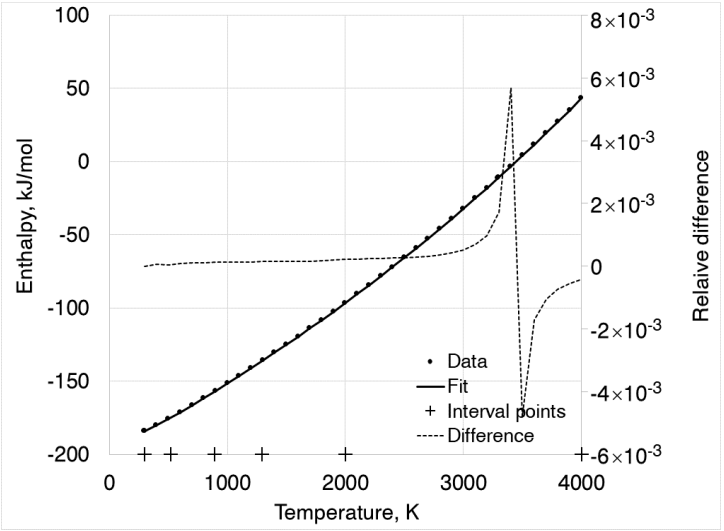


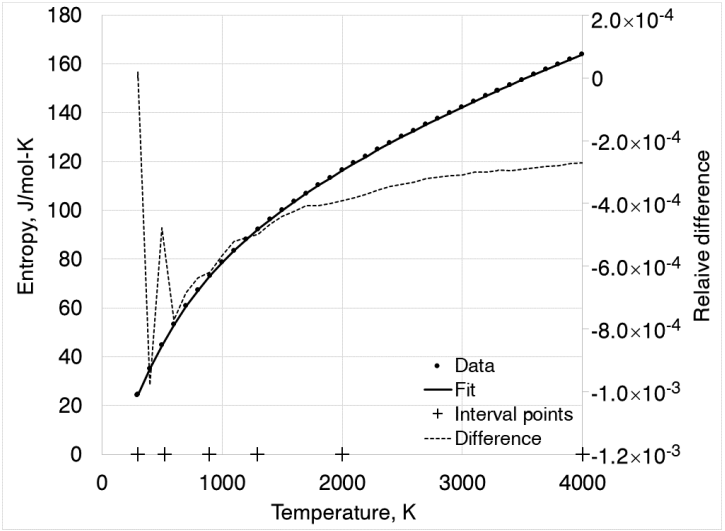


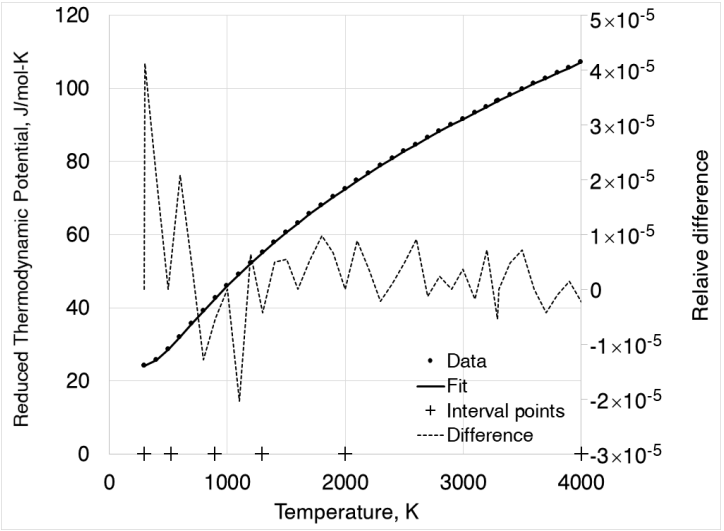
Titanium Carbide (TiC) – C1Ti1(cr)

MW = 59.89 g/mol – $\Delta_f H^o$(298.15 K) = -184.096 kJ/mol – S^o(298.15 K) = 24.229 J/mol-K – p^o = 0.1 MPa



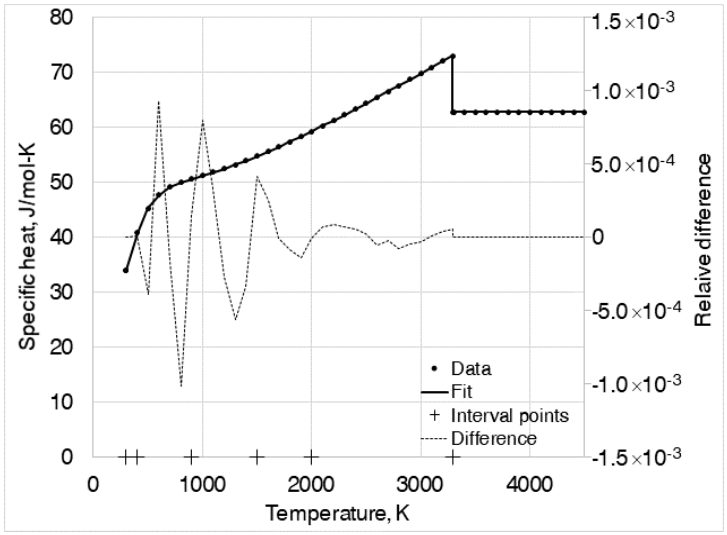


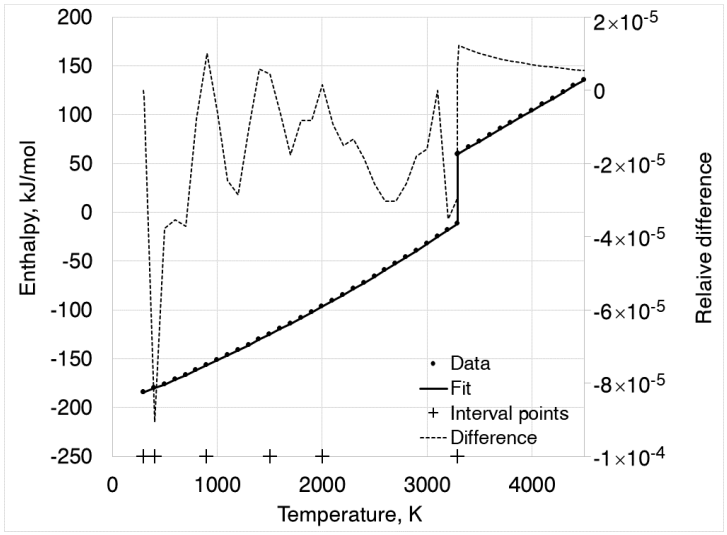


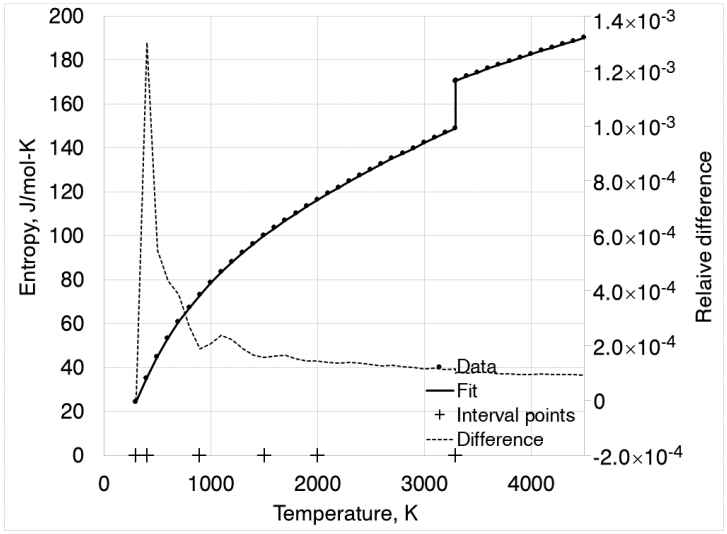


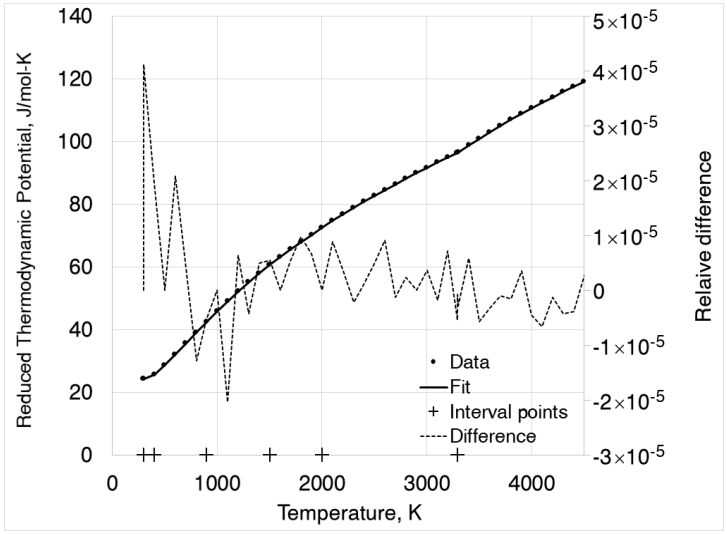
Titanium Carbide (TiC) – C1Ti1(cr,l)

MW = 59.89 g/mol – $\Delta_f H^o$(298.15 K) = -184.096 kJ/mol – S^o(298.15 K) = 24.229 J/mol-K – p^o = 0.1 MPa



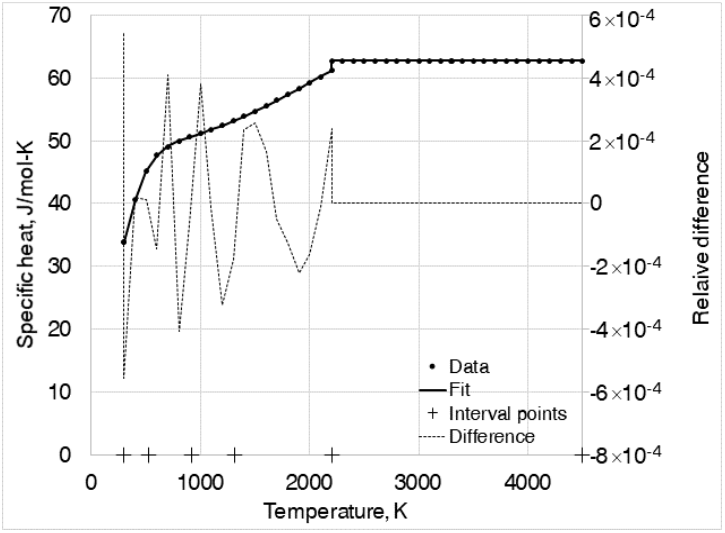


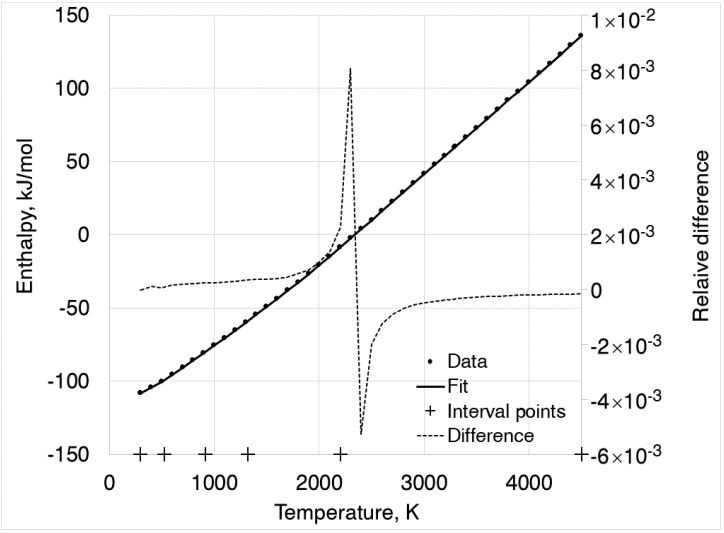


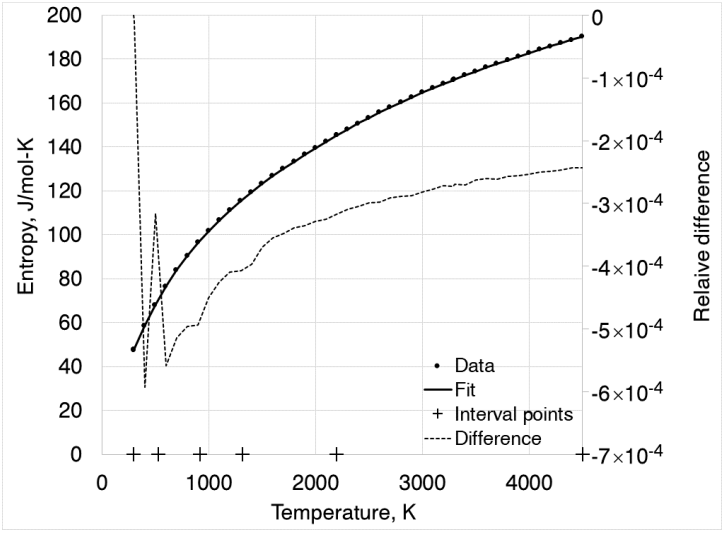


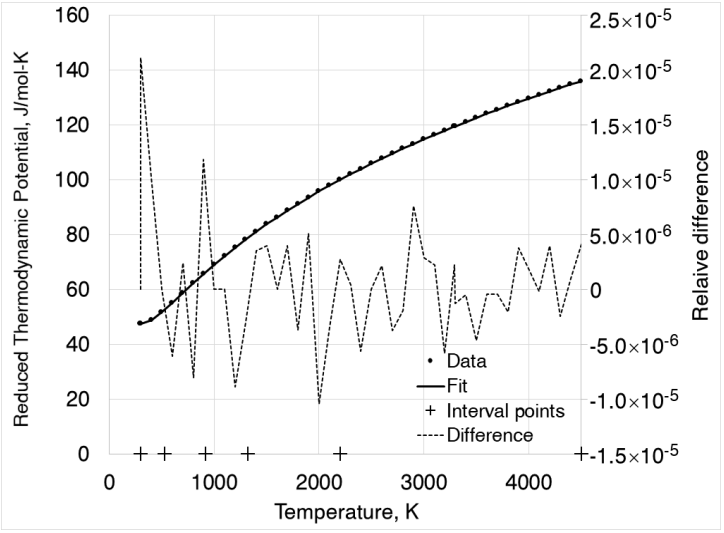
Titanium Carbide (TiC) – C1Ti1(l)

MW = 59.89 g/mol – $\Delta_f H^o$(298.15 K) = -108.334 kJ/mol – S^o(298.15 K) = 47.401 J/mol-K – p^o = 0.1 MPa



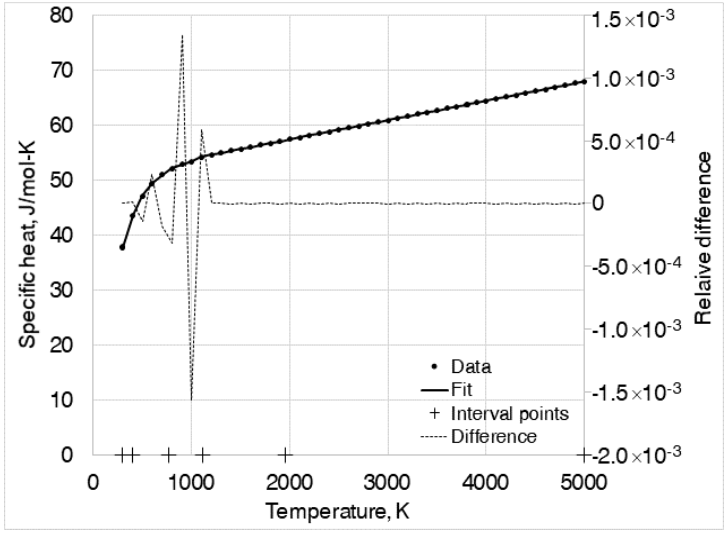


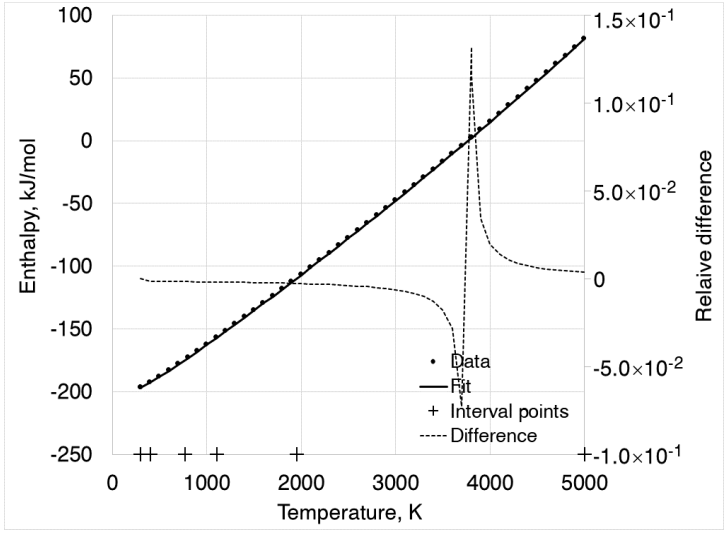


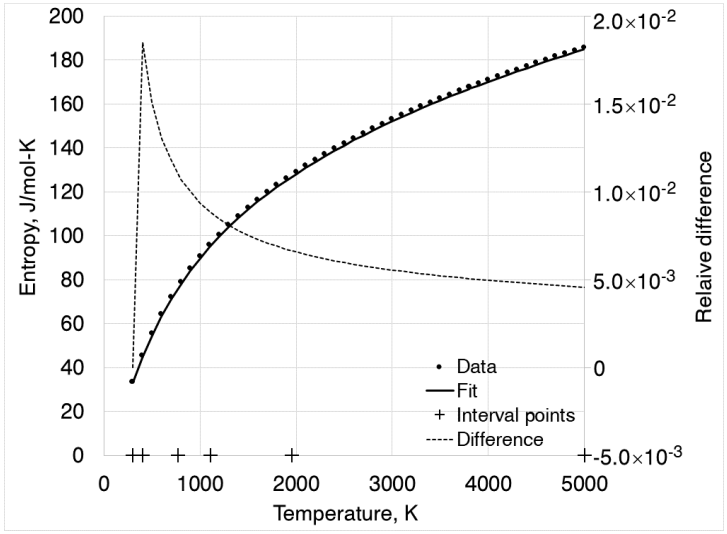


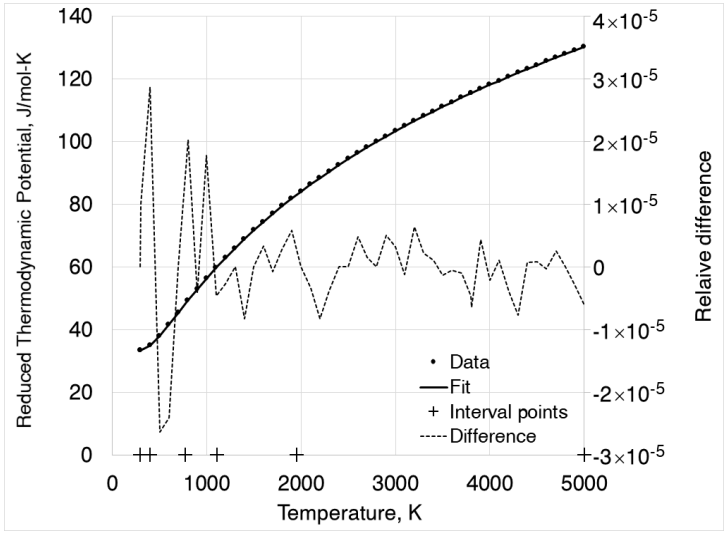
Zirconium Carbide (ZrC) – C1Zr1(cr)

MW = 103.23 g/mol – $\Delta_f H^o$(298.15 K) = -196.648 kJ/mol – S^o(298.15 K) = 33.321 J/mol-K – p^o = 0.1 MPa



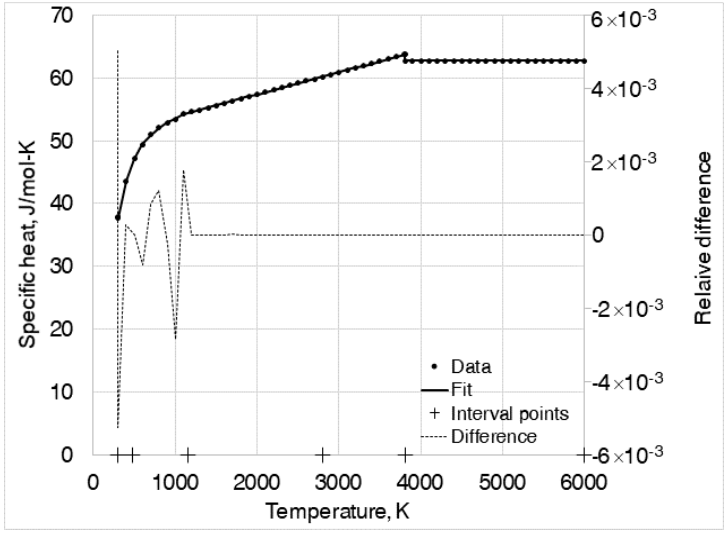


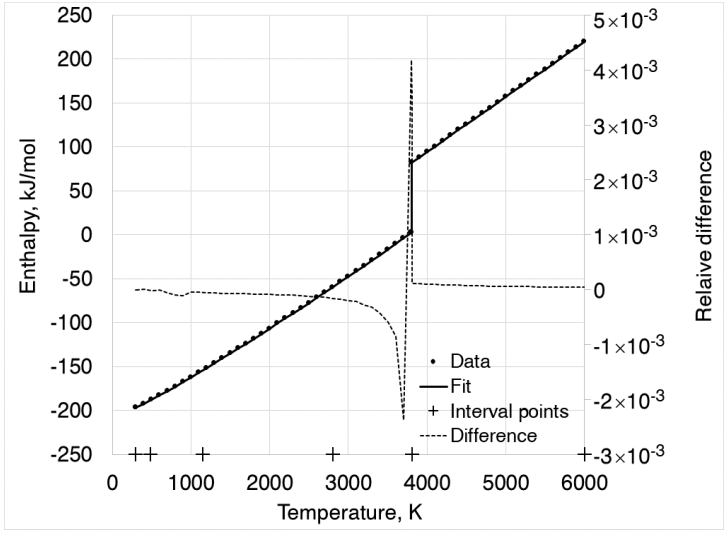


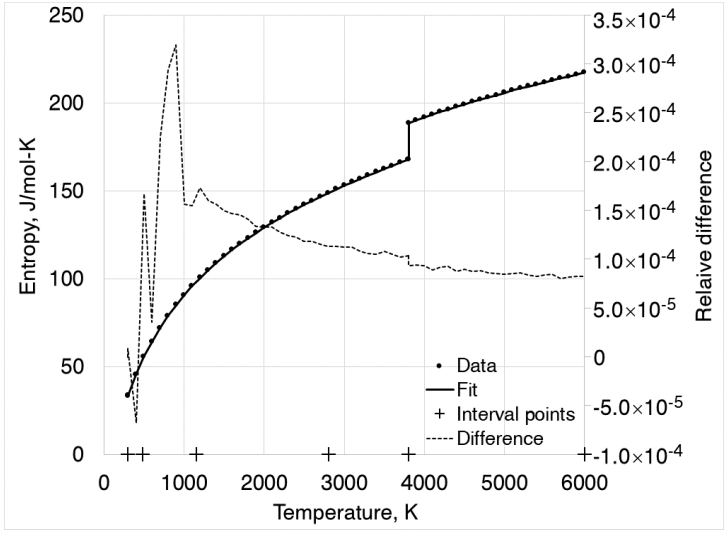


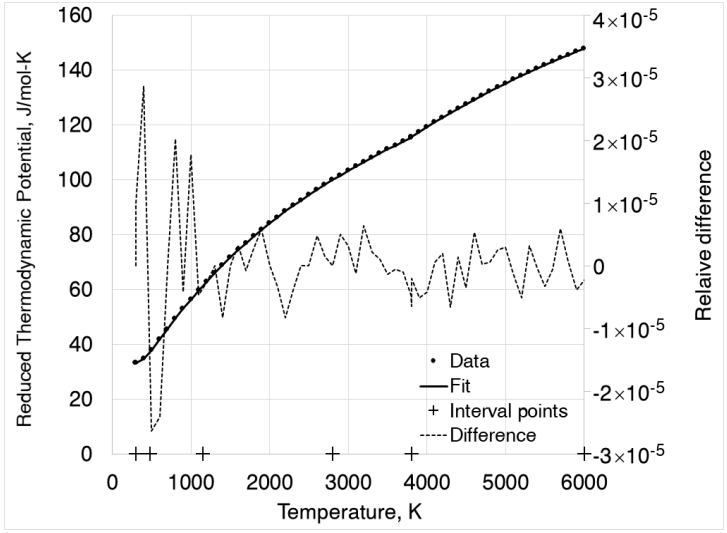
Zirconium Carbide (ZrC) – C1Zr1(cr,l)

MW = 103.23 g/mol – $\Delta_f H^o$(298.15 K) = -196.648 kJ/mol – S^o(298.15 K) = 33.321 J/mol-K – p^o = 0.1 MPa



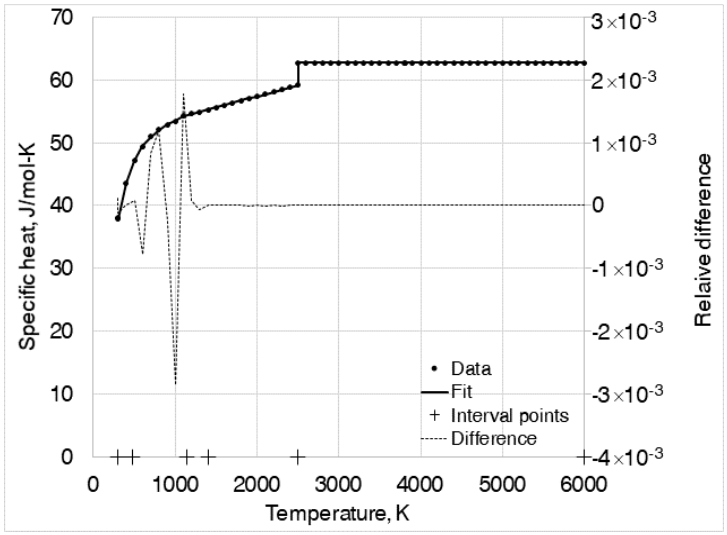


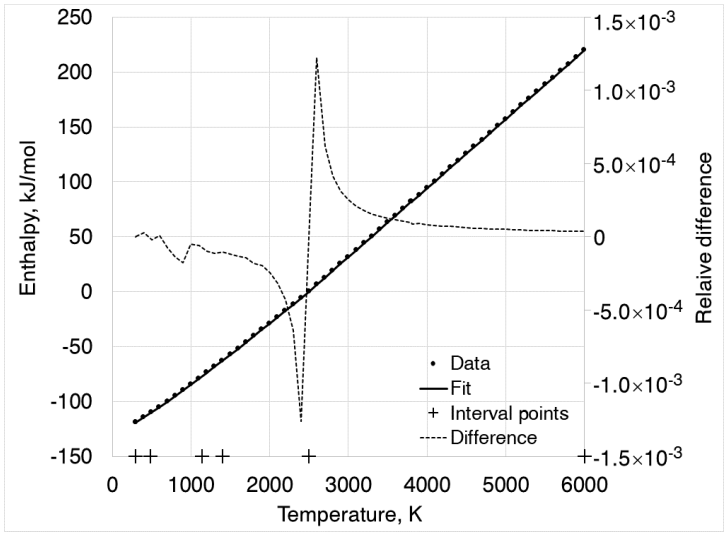


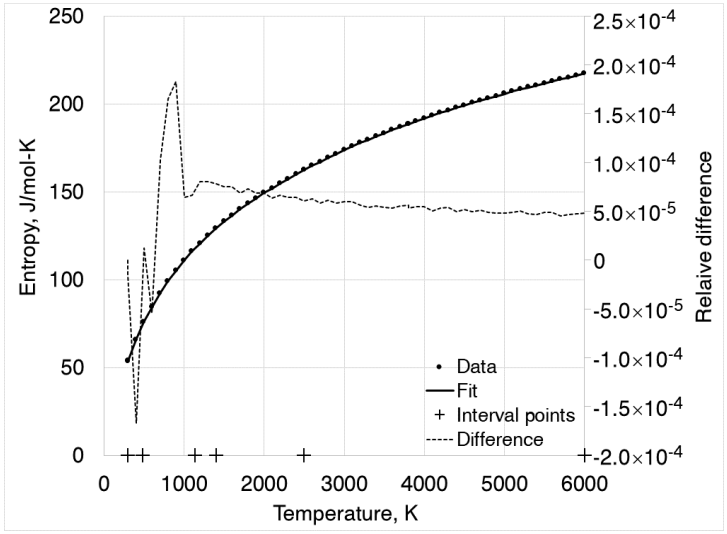


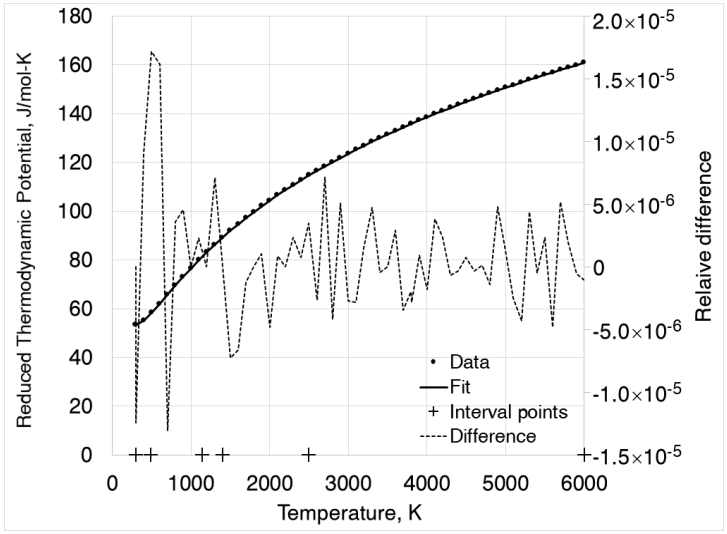
Zirconium Carbide (ZrC) – C1Zr1(l)

MW = 103.23 g/mol – $\Delta_f H^o$(298.15 K) = -118.877 kJ/mol – S^o(298.15 K) = 53.577 J/mol-K – p^o = 0.1 MPa



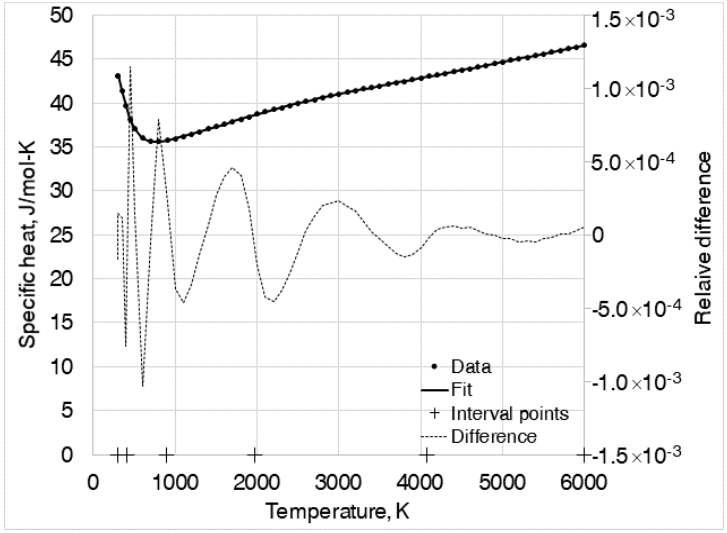


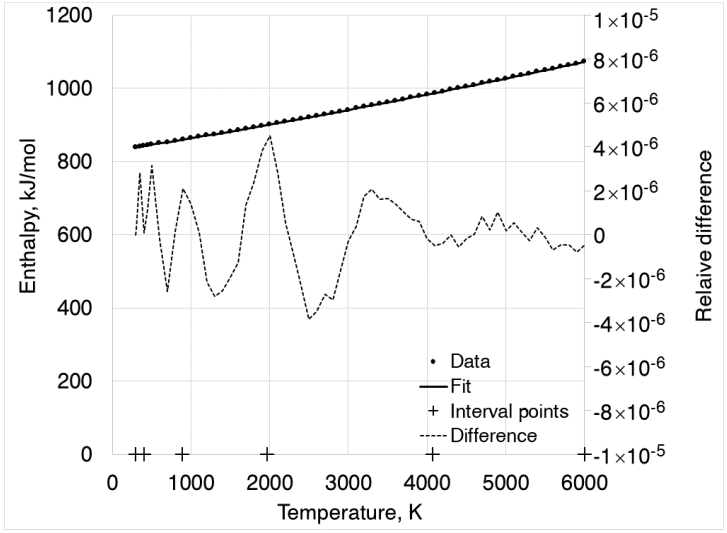


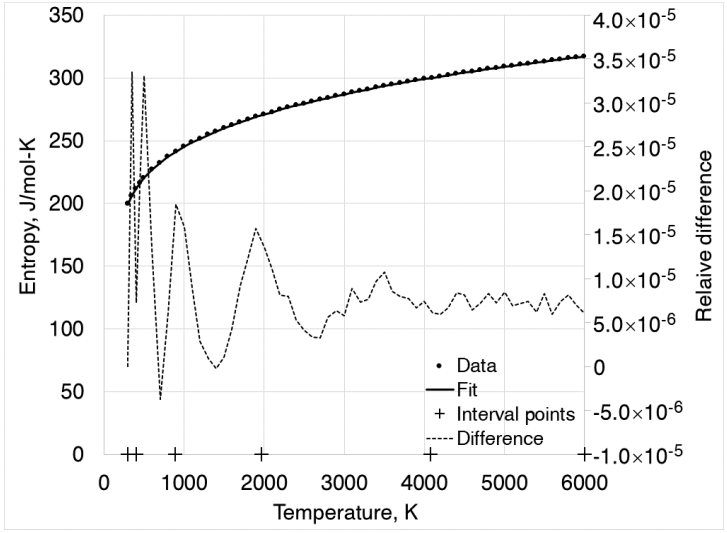


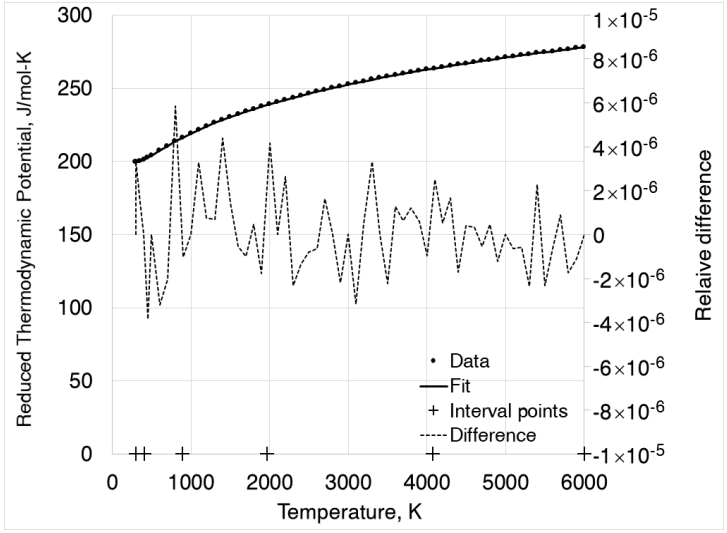
Carbon (C_2) – C2(g)

MW = 24.022 g/mol – $\Delta_f H^o$(298.15 K) = 837.737 kJ/mol – S^o(298.15 K) = 199.382 J/mol-K – p^o = 0.1 MPa



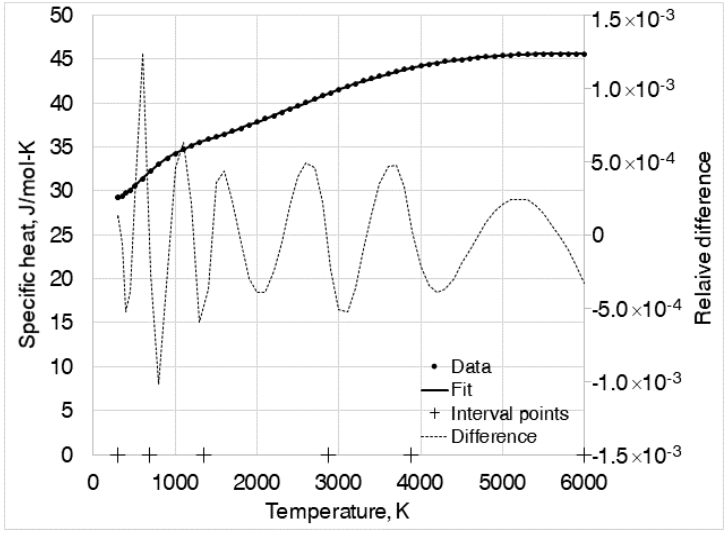


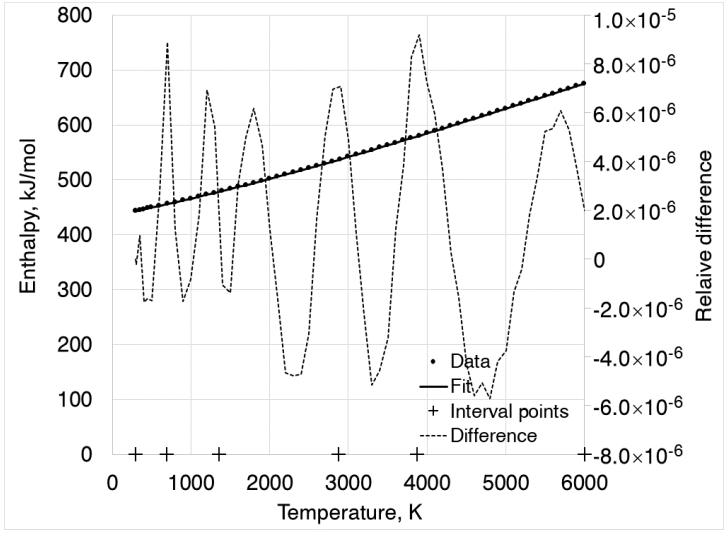


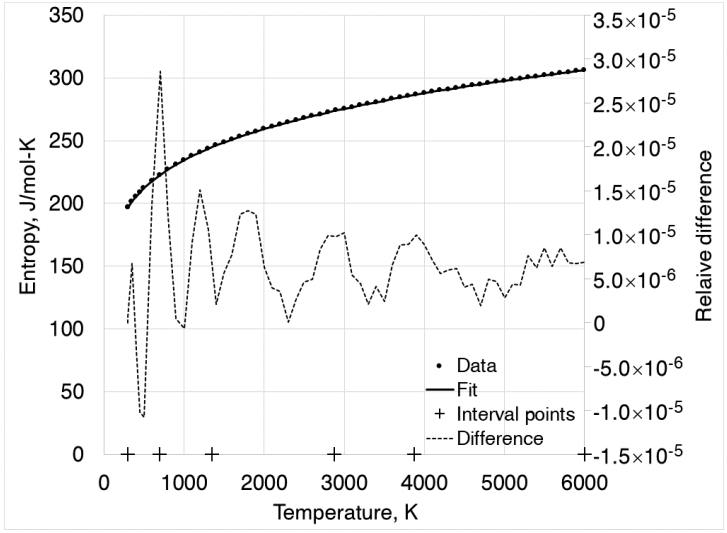


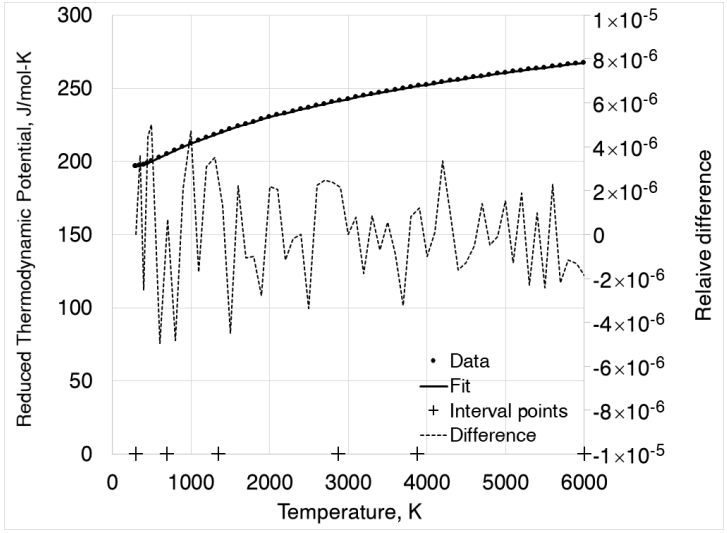
Carbon, Ion (C_2^-) – C2-(g)

MW = 24.023 g/mol – $\Delta_f H^o$(298.15 K) = 443.504 kJ/mol – S^o(298.15 K) = 196.597 J/mol-K – p^o = 0.1 MPa



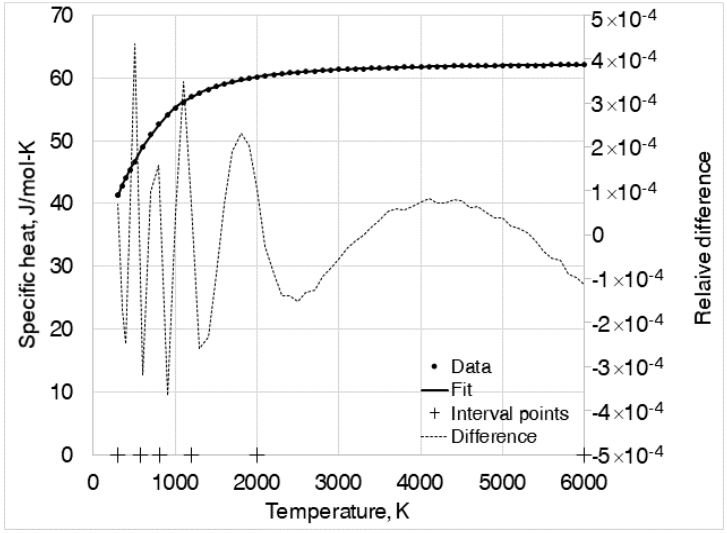


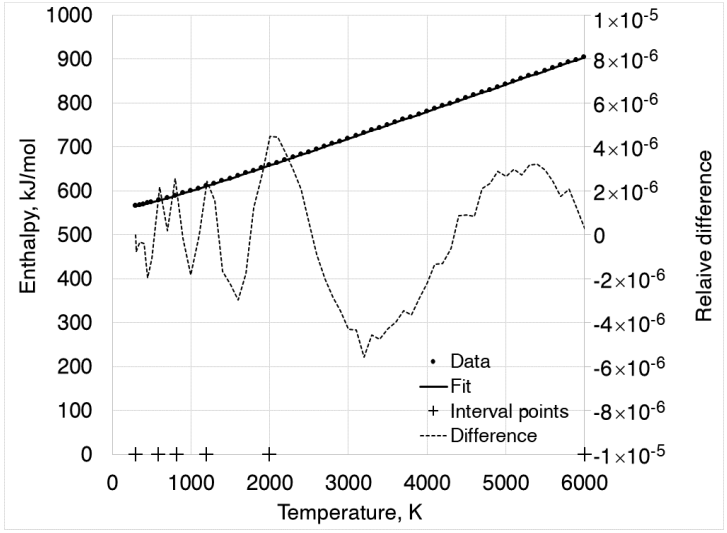


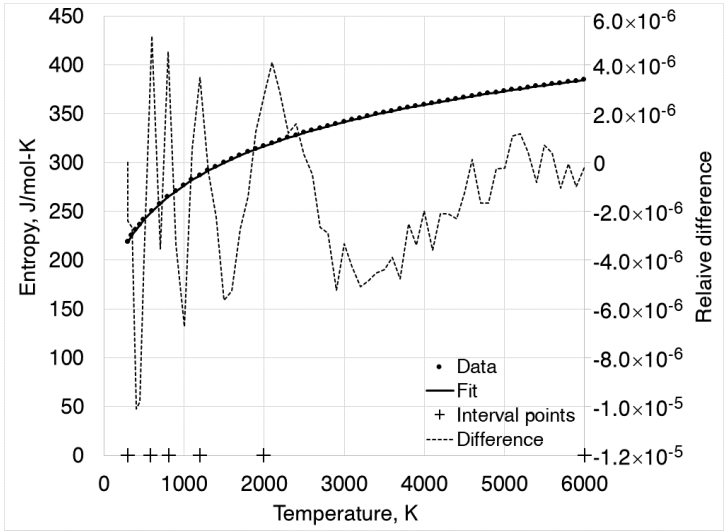


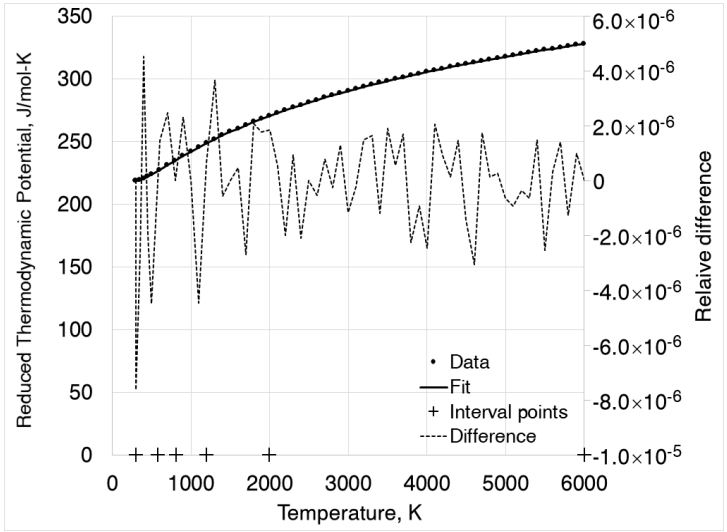
Beryllium Carbide (BeC_2) – C2Be1(g)

MW = 33.034 g/mol – $\Delta_f H^o$(298.15 K) = 564.84 kJ/mol – S^o(298.15 K) = 218.643 J/mol-K – p^o = 0.1 MPa



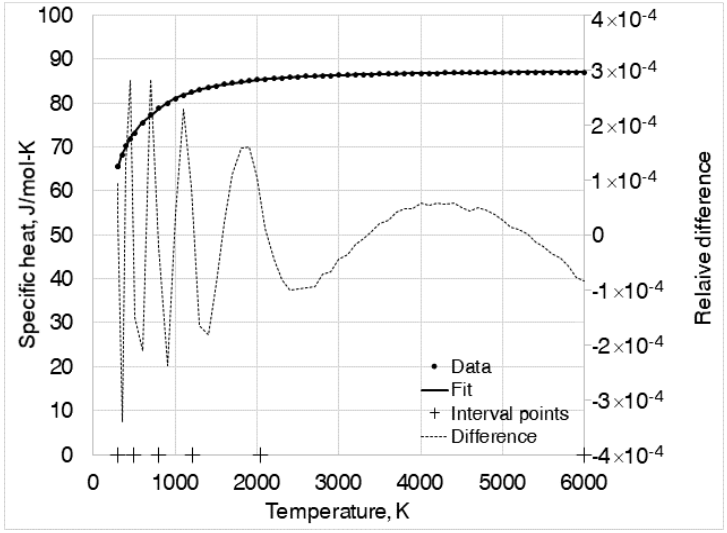


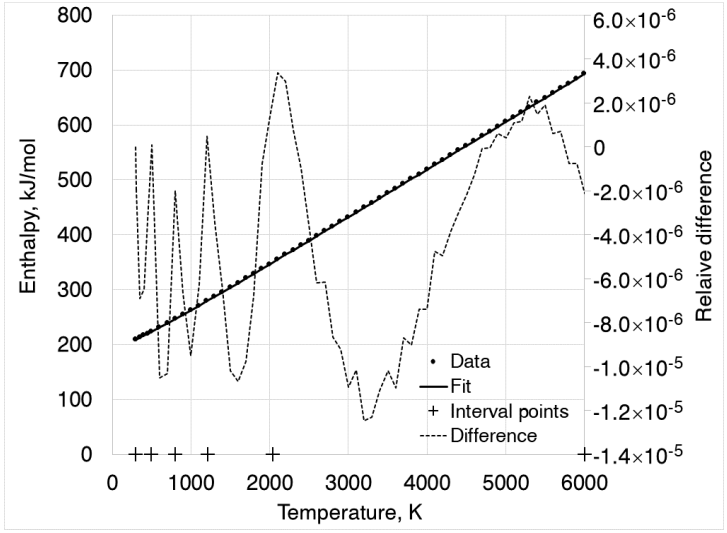


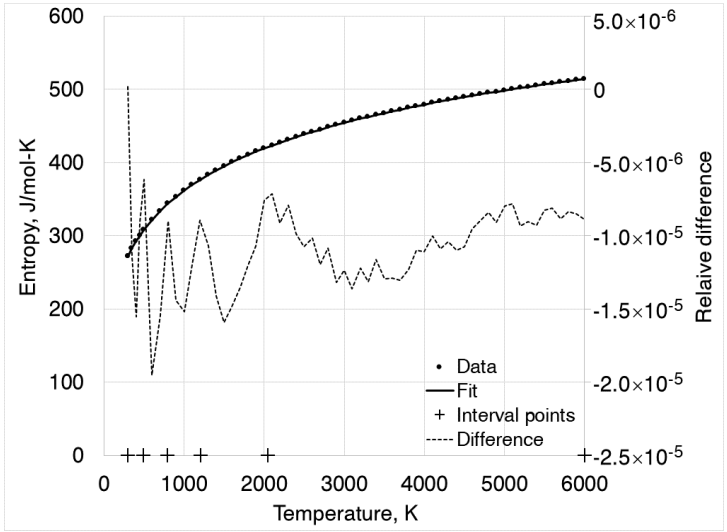


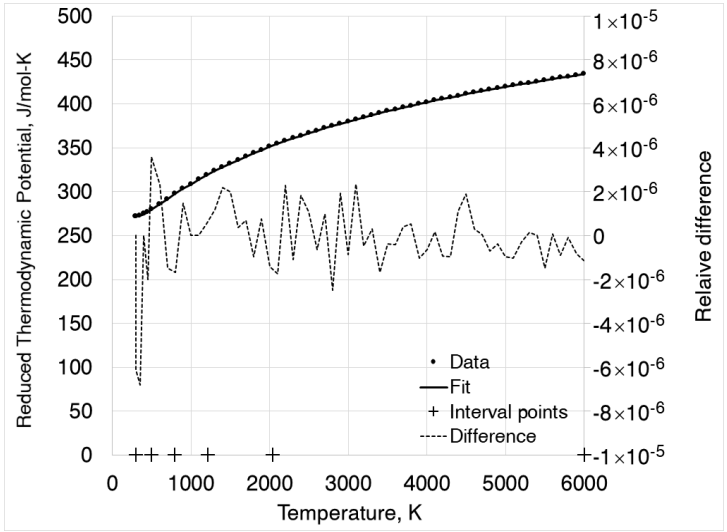
Dichloroethyne (C_2Cl_2) – C2Cl2(g)

MW = 94.928 g/mol – $\Delta_f H^o$(298.15 K) = 209.618 kJ/mol – S^o(298.15 K) = 272.027 J/mol-K – p^o = 0.1 MPa



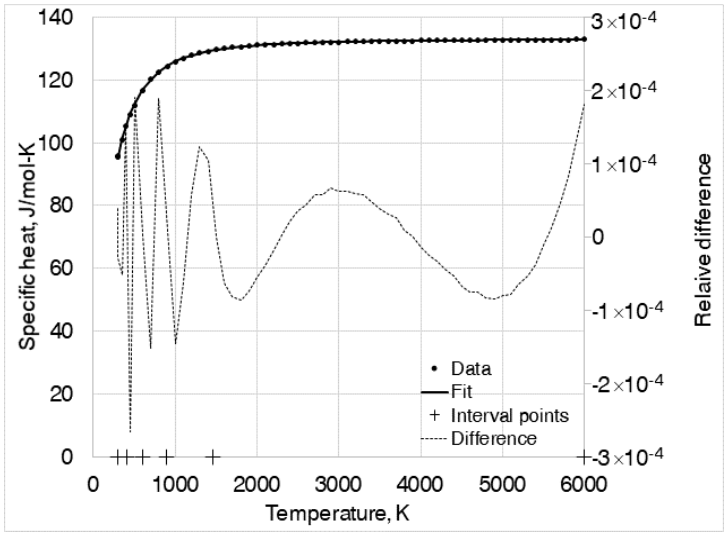


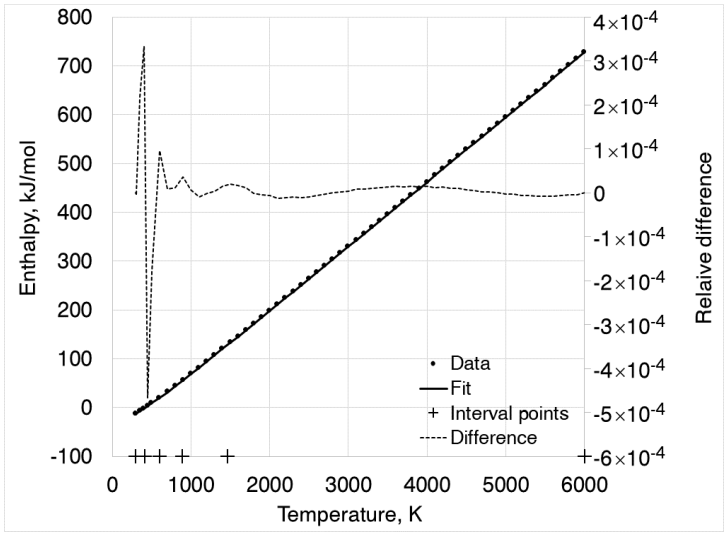


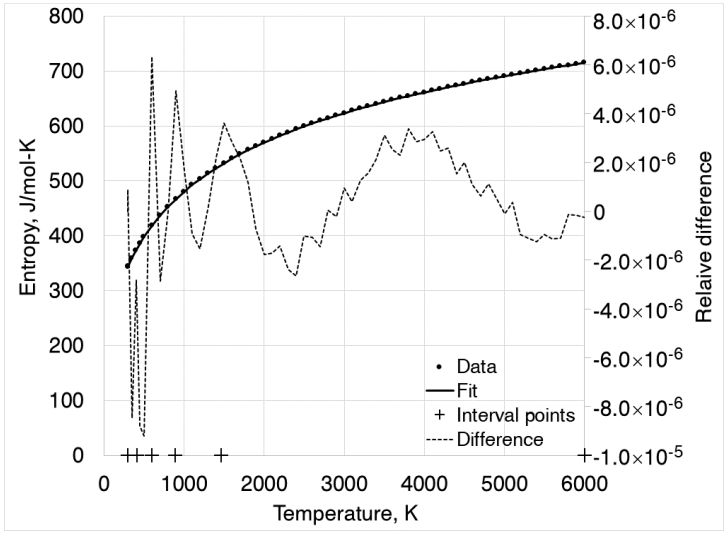


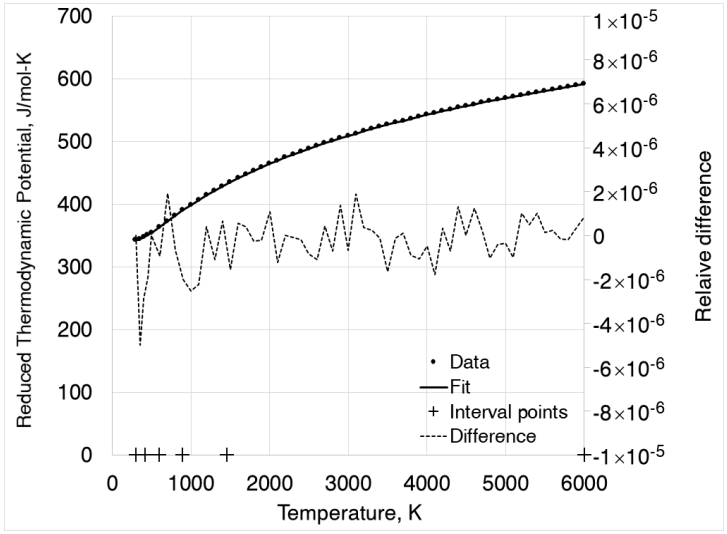
Tetrachloroethene (C_2Cl_4) – C2Cl4(g)

MW = 165.834 g/mol – $\Delta_f H^o$(298.15 K) = -12.426 kJ/mol – S^o(298.15 K) = 343.421 J/mol-K – p^o = 0.1 MPa



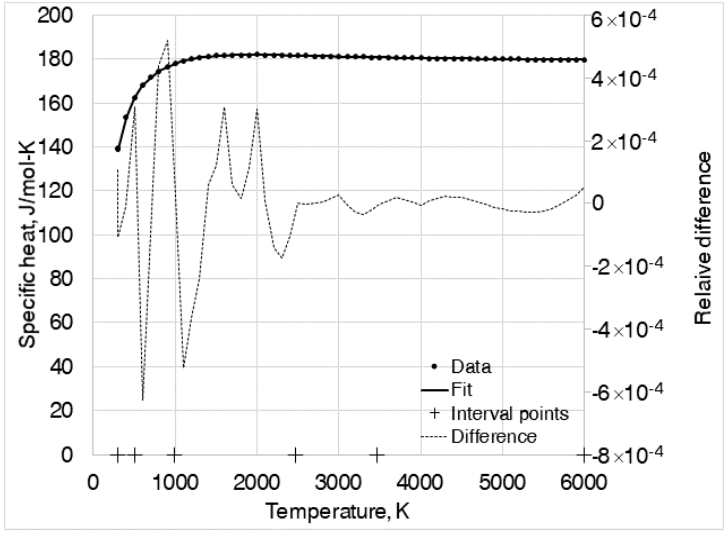


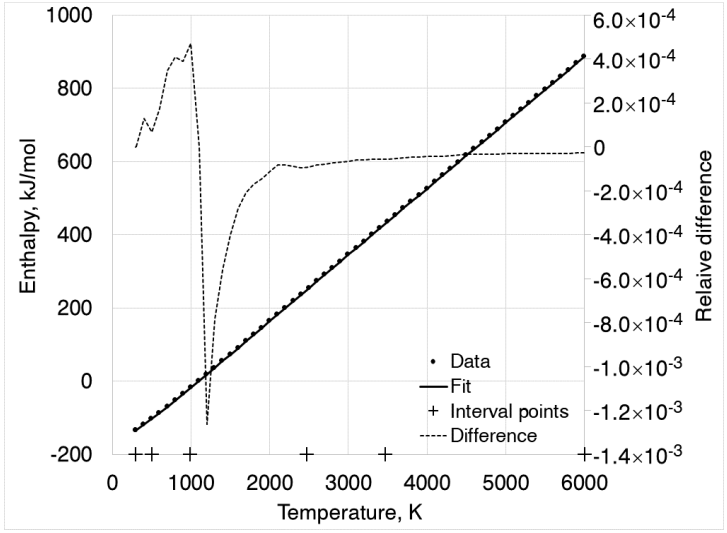


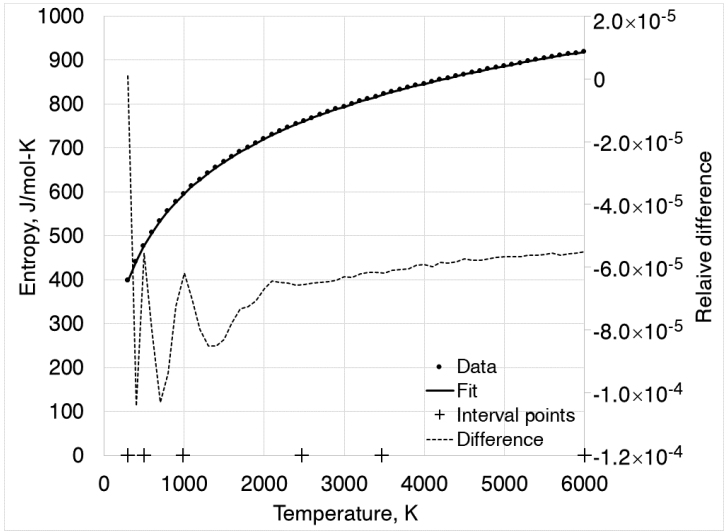


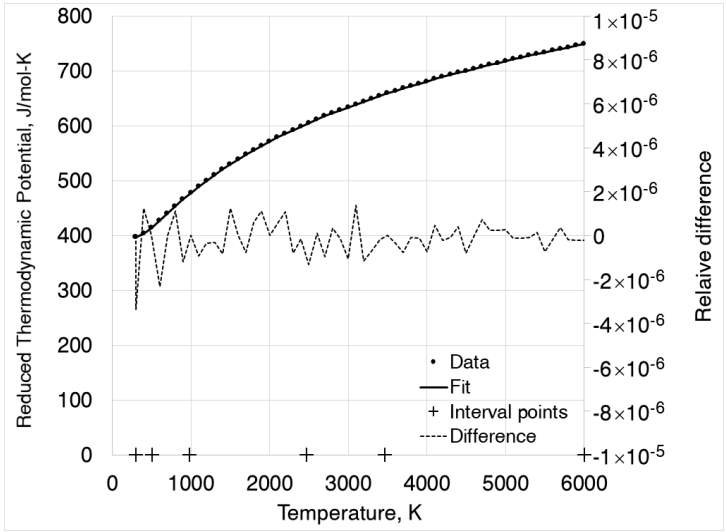
Hexachloroethane (C_2Cl_6) – C2Cl6(g)

MW = 236.74 g/mol – $\Delta_f H^o$(298.15 K) = -134.223 kJ/mol – S^o(298.15 K) = 397.88 J/mol-K – p^o = 0.1 MPa



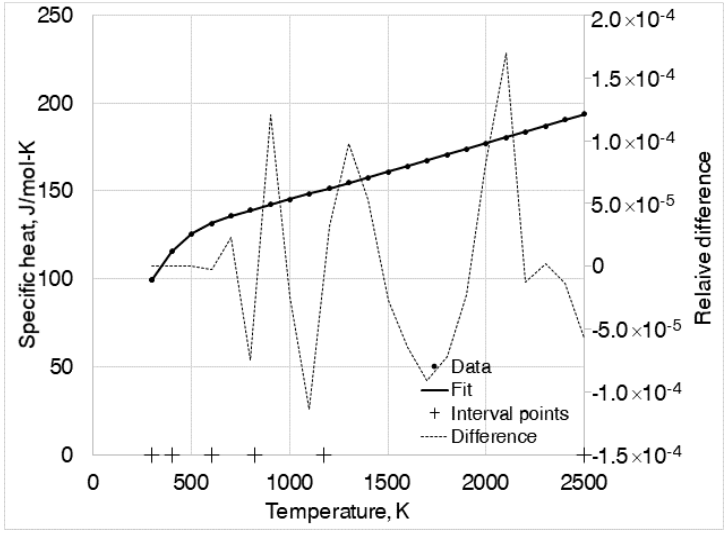


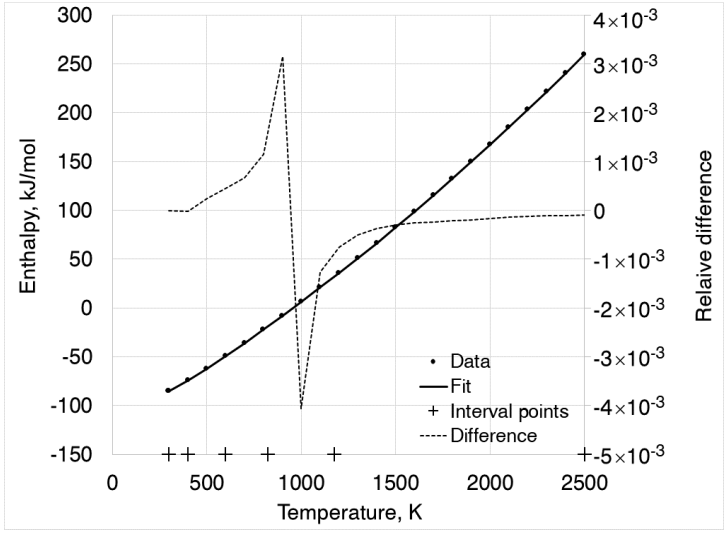


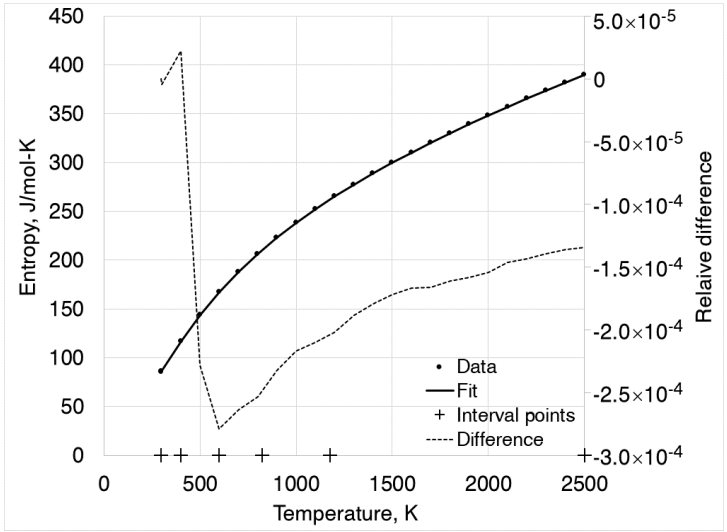


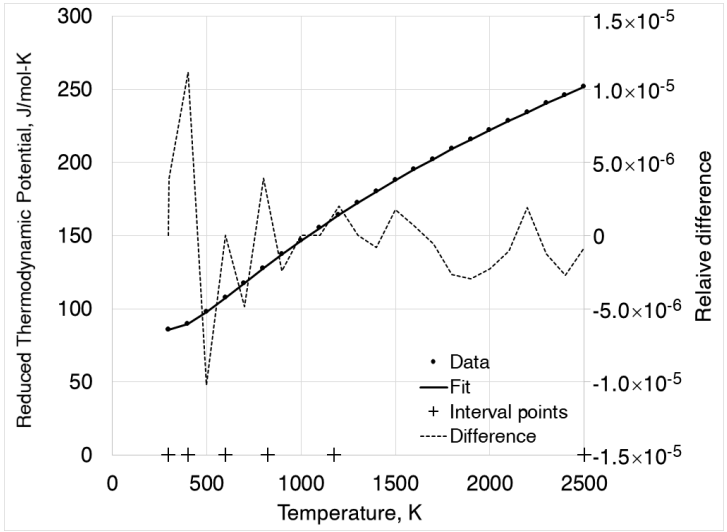
Chromium Carbide (Cr_3C_2) – C2Cr3(cr)

MW = 180.01 g/mol – $\Delta_f H^o$(298.15 K) = -85.354 kJ/mol – S^o(298.15 K) = 85.434 J/mol-K – p^o = 0.1 MPa



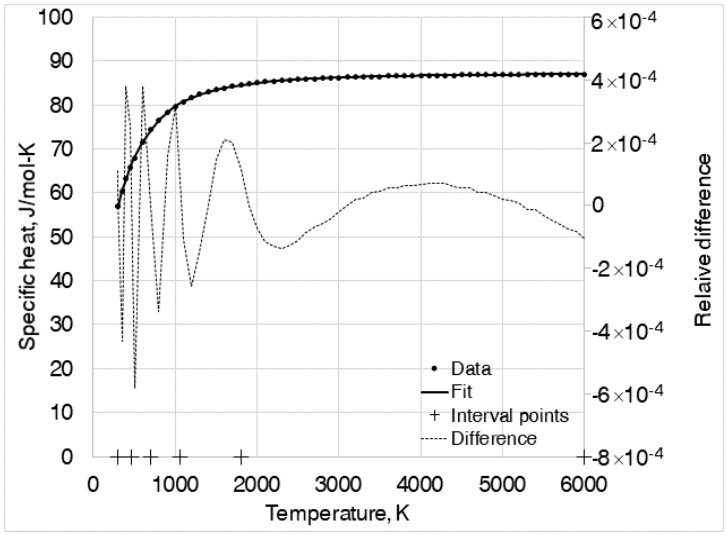


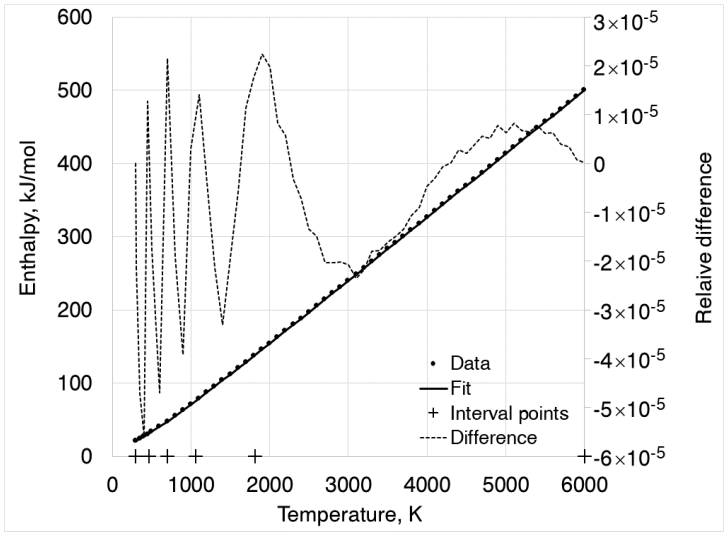


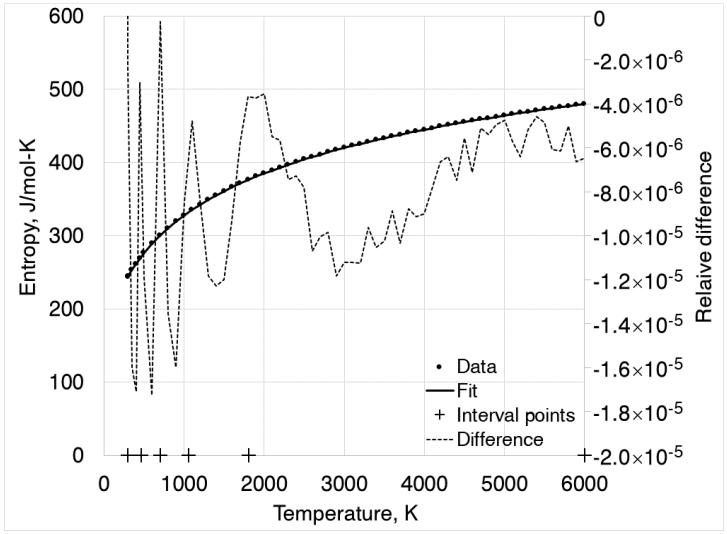


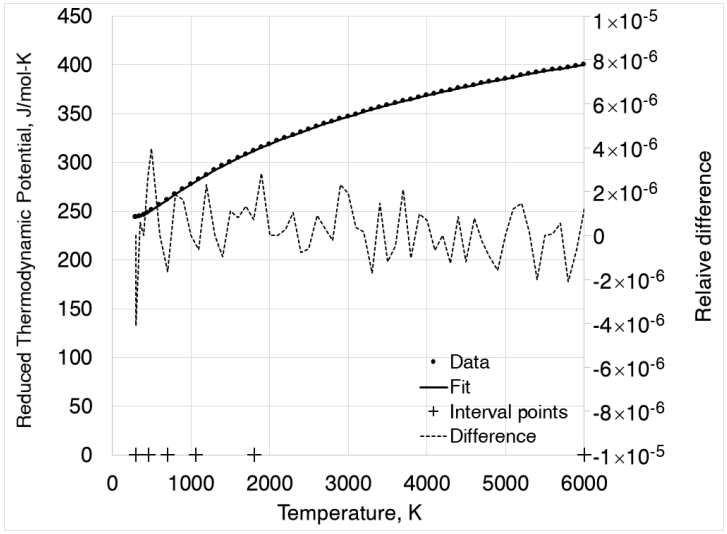
Difluoroethyne (C_2F_2) – C2F2(g)

MW = 62.019 g/mol – $\Delta_f H^o$(298.15 K) = 20.92 kJ/mol – S^o(298.15 K) = 244.06 J/mol-K – p^o = 0.1 MPa



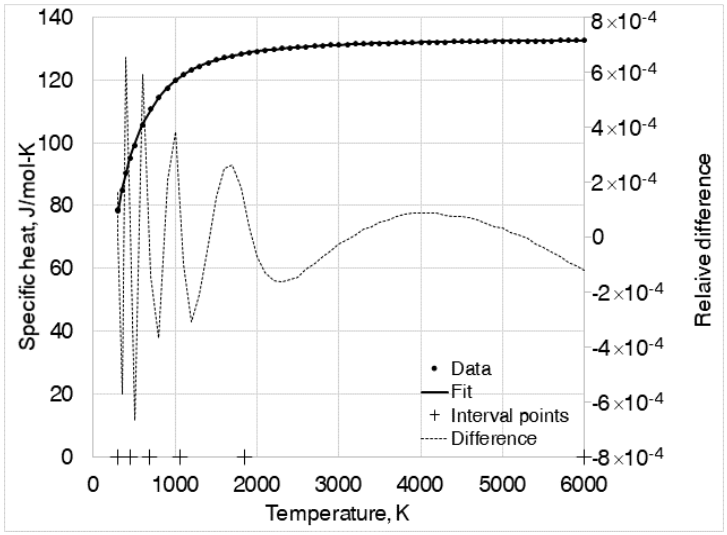


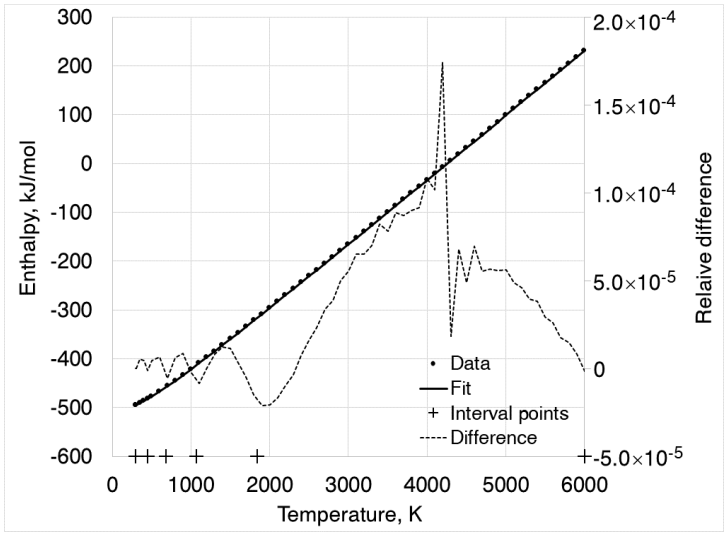


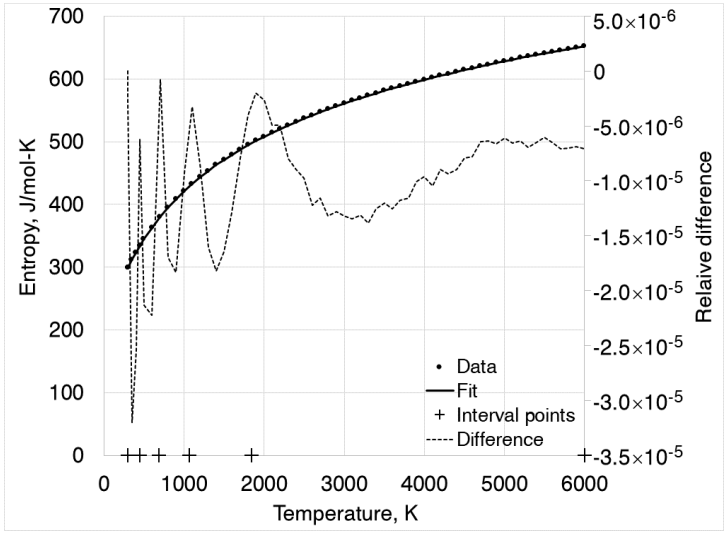


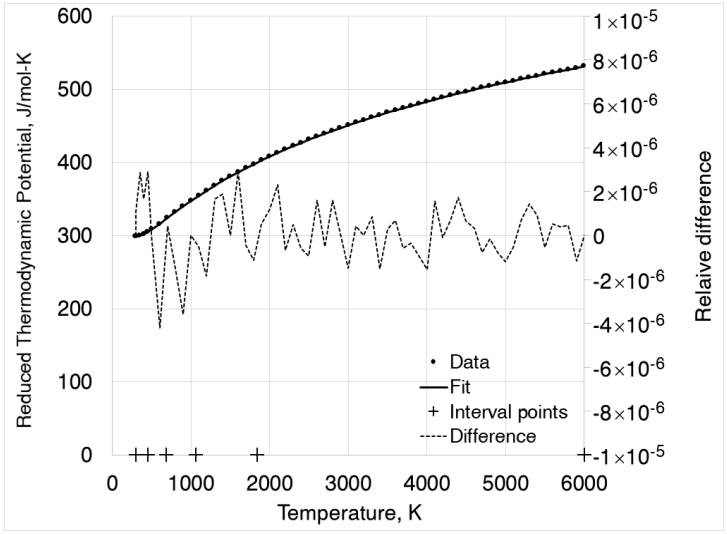
Trifluoroacetonitrile (CF_3CN) – C2F3N1(g)

MW = 95.024 g/mol – $\Delta_f H^o$(298.15 K) = -495.386 kJ/mol – S^o(298.15 K) = 298.533 J/mol-K – p^o = 0.1 MPa



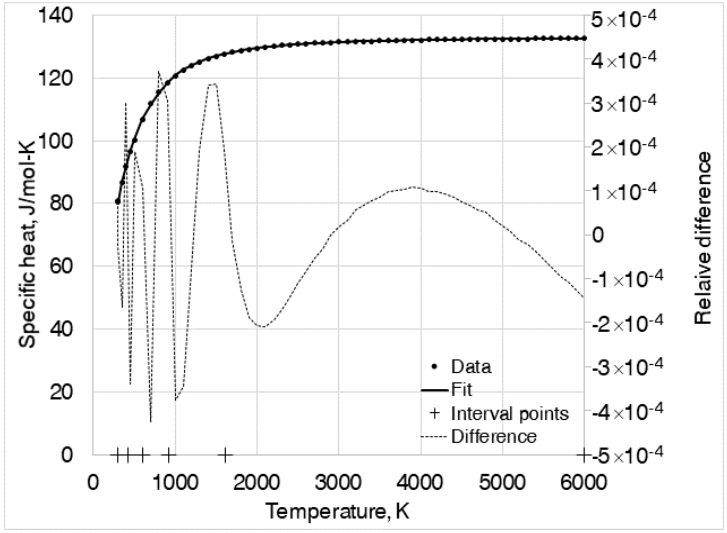


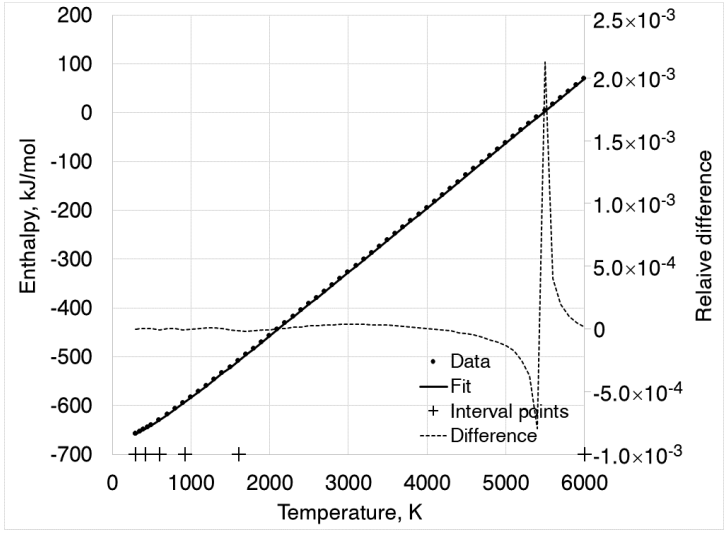


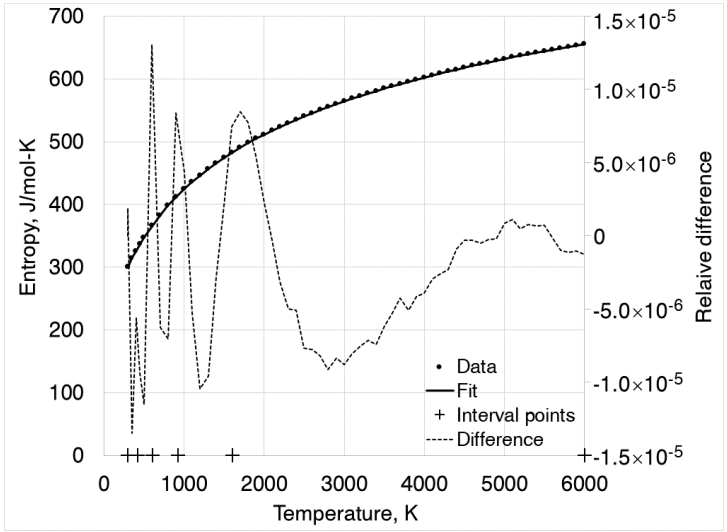


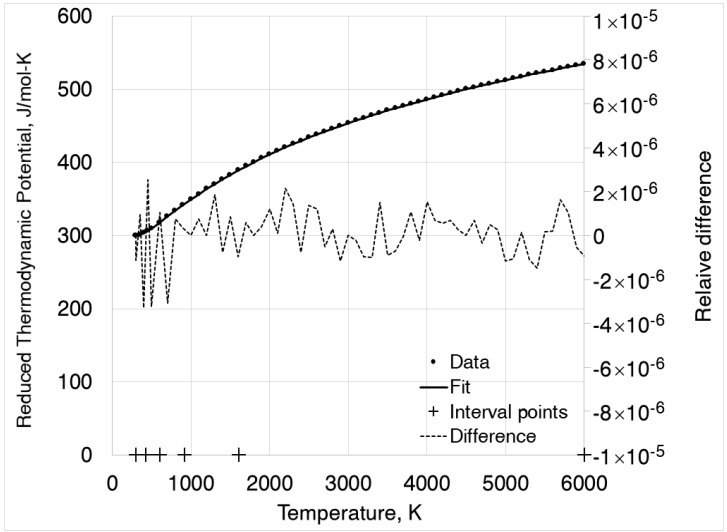
Tetrafluoroethene (C_2F_4) – C2F4(g)

MW = 100.016 g/mol – $\Delta_f H^o$(298.15 K) = -658.562 kJ/mol – S^o(298.15 K) = 300.015 J/mol-K – p^o = 0.1 MPa



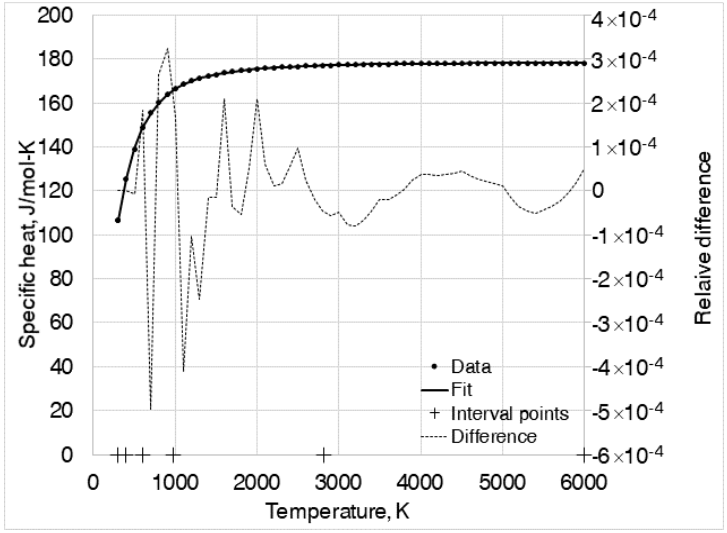


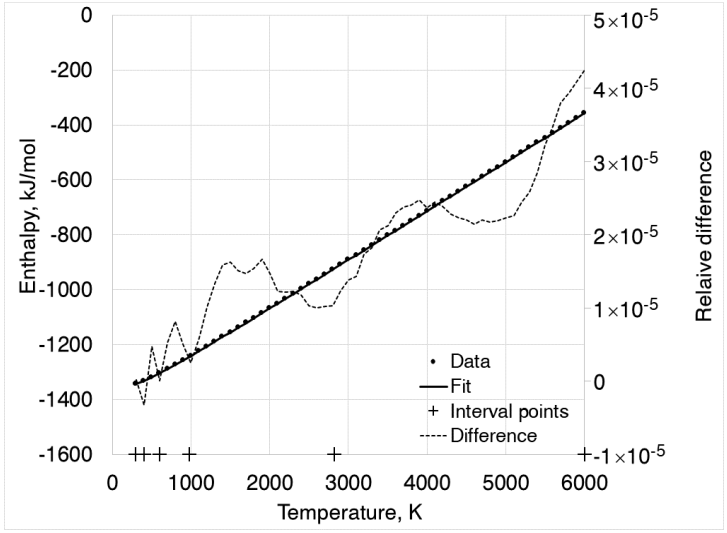


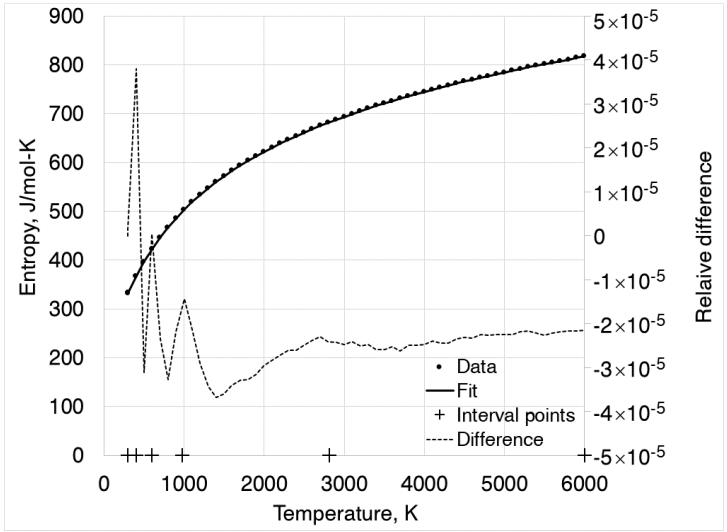


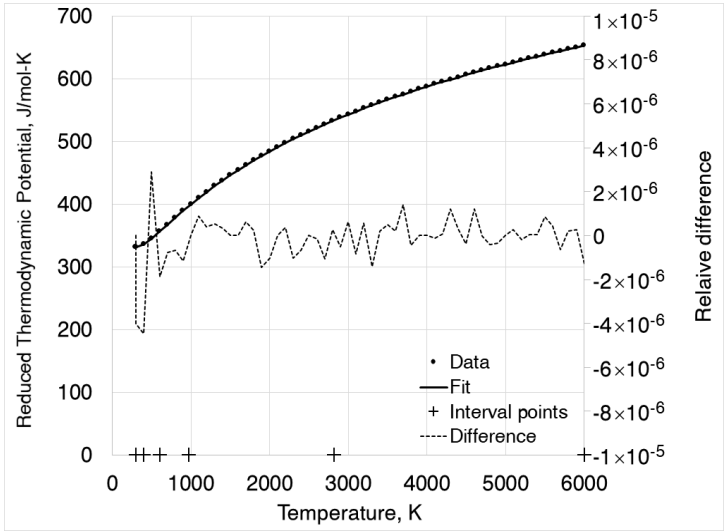
Hexafluoroethane (C_2F_6) – C2F6(g)

MW = 138.012 g/mol – $\Delta_f H^o$(298.15 K) = -1343.901 kJ/mol – S^o(298.15 K) = 332.185 J/mol-K – p^o = 0.1 MPa



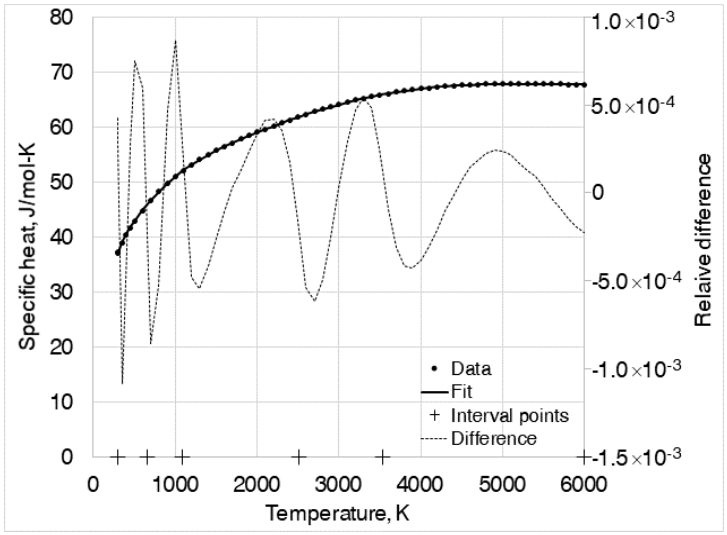


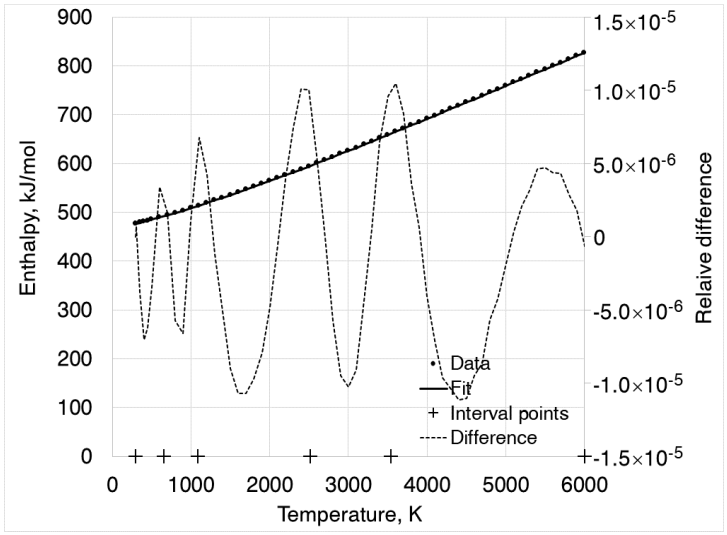


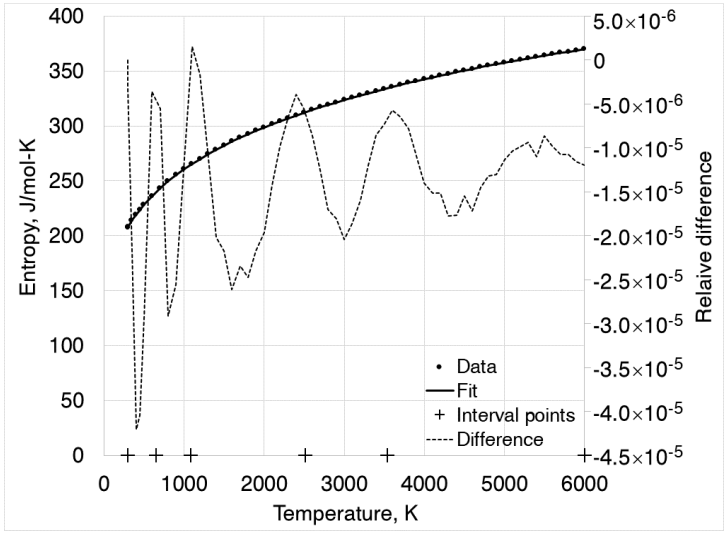


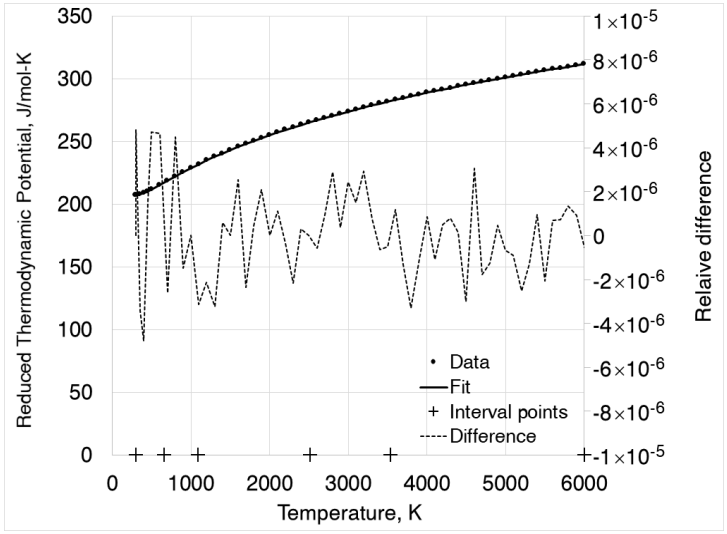
Ethynyl (C_2H) – C2H1(g)

MW = 25.03 g/mol – $\Delta_f H^o$(298.15 K) = 476.976 kJ/mol – S^o(298.15 K) = 207.444 J/mol-K – p^o = 0.1 MPa



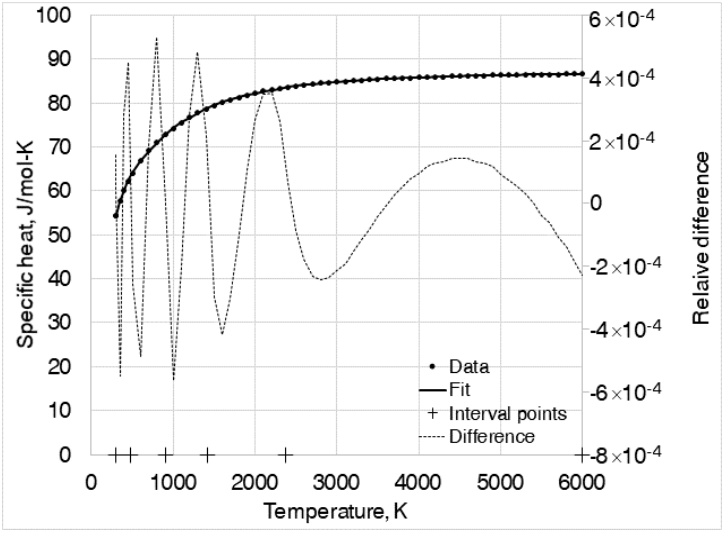


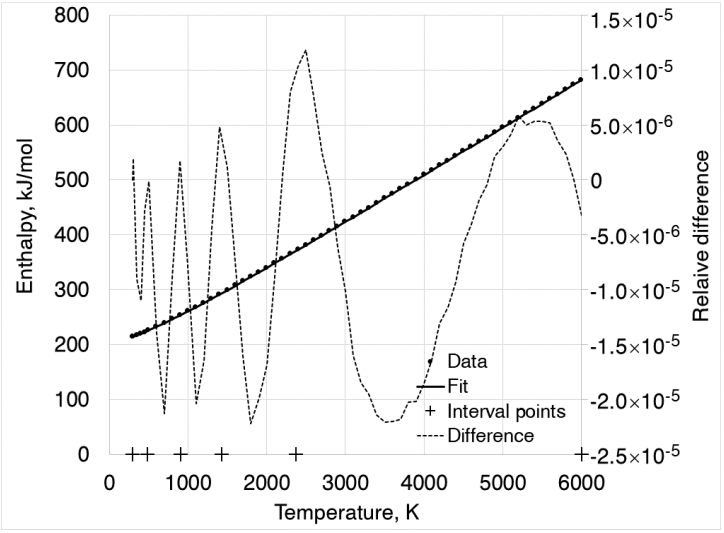


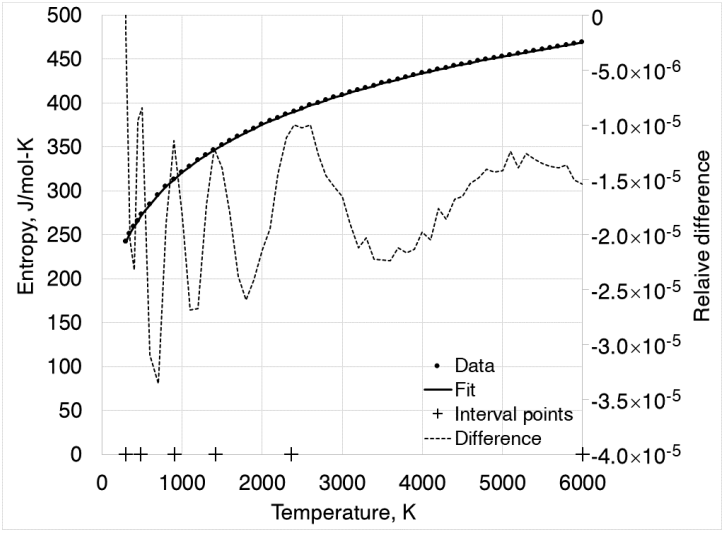


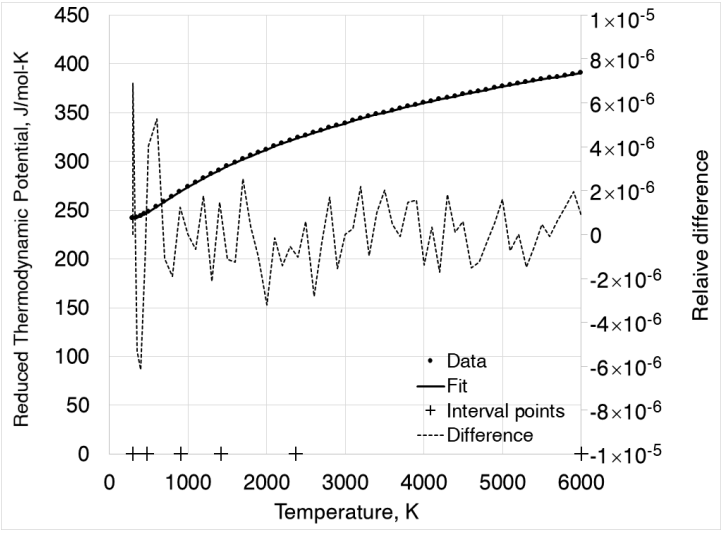
Chloroethyne (C_2HCl) – C2H1Cl1(g)

MW = 60.483 g/mol – $\Delta_f H^o$(298.15 K) = 213.802 kJ/mol – S^o(298.15 K) = 241.996 J/mol-K – p^o = 0.1 MPa



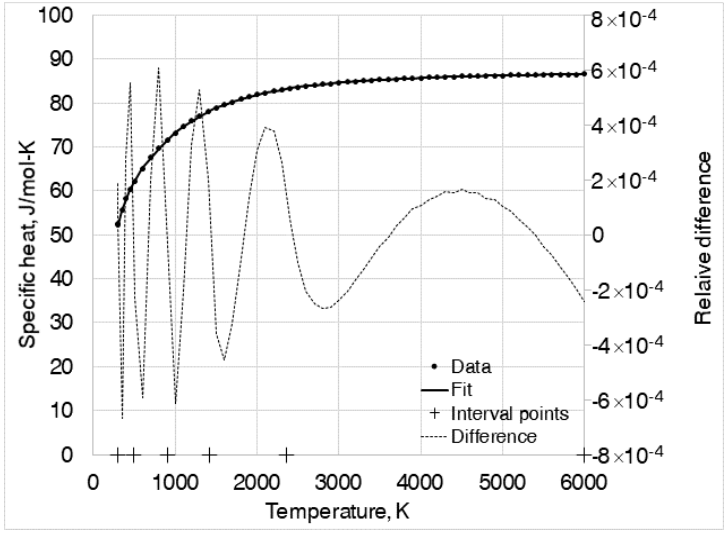


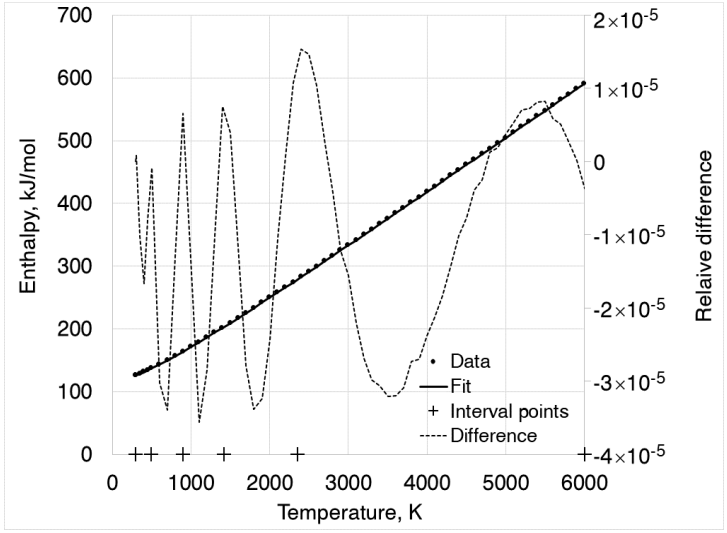


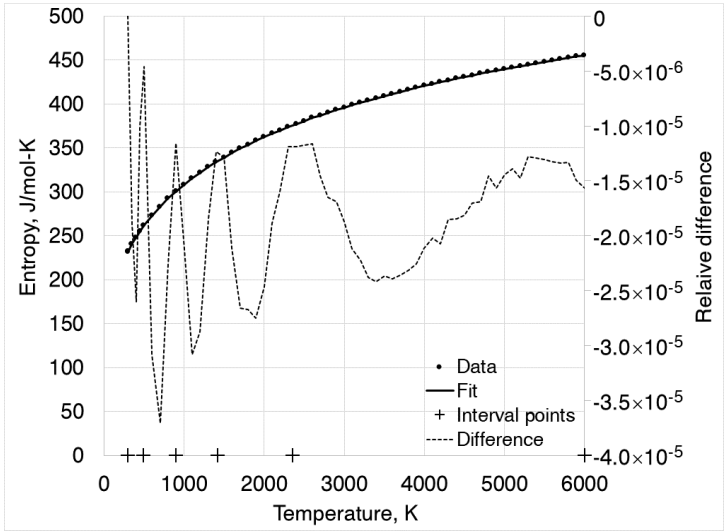


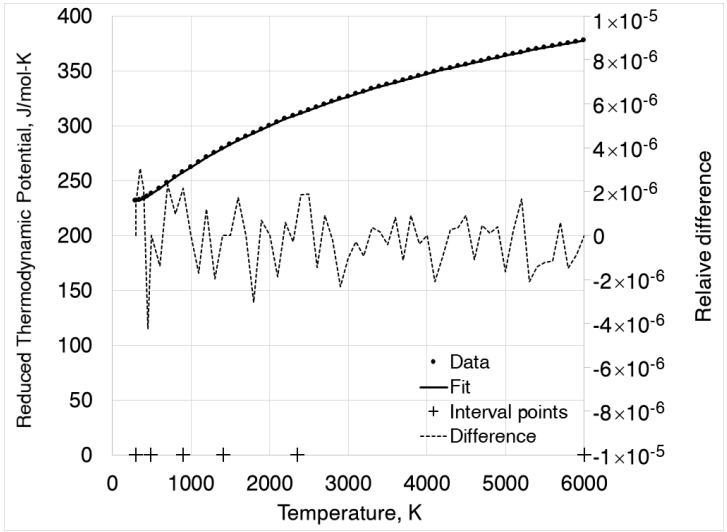
Fluoroethyne (C_2HF) – C2H1F1(g)

MW = 44.028 g/mol – $\Delta_f H^o$(298.15 K) = 125.52 kJ/mol – S^o(298.15 K) = 231.611 J/mol-K – p^o = 0.1 MPa



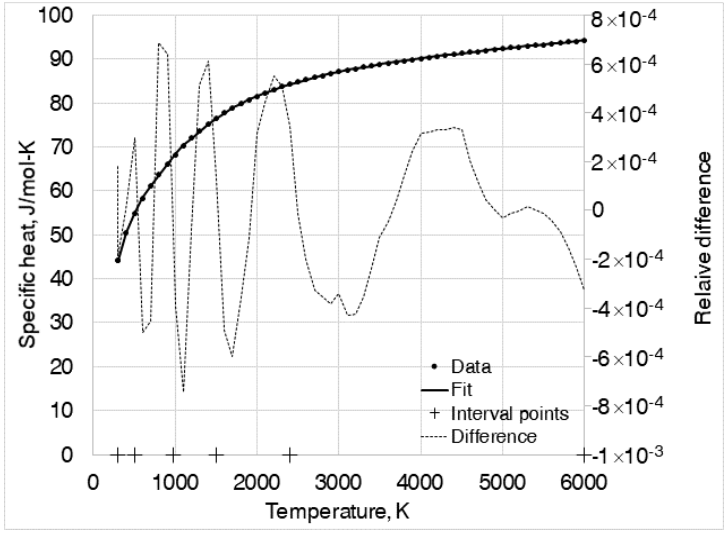


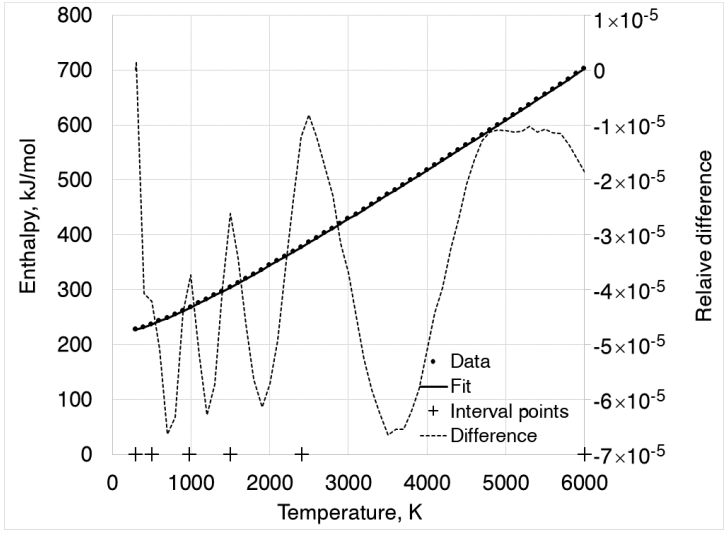


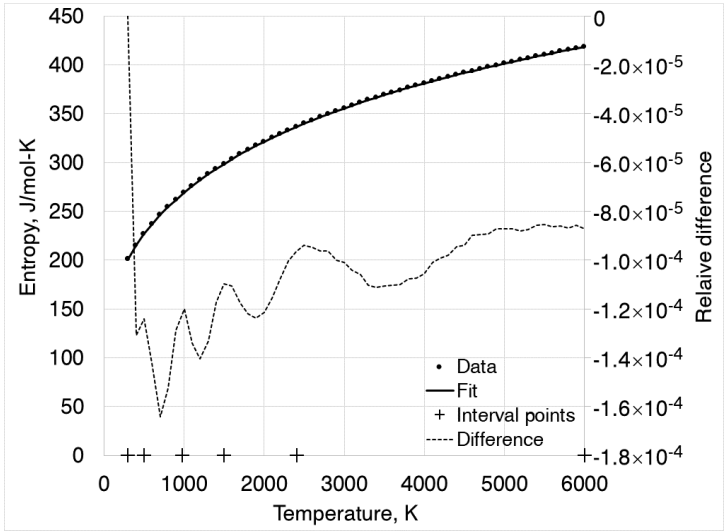


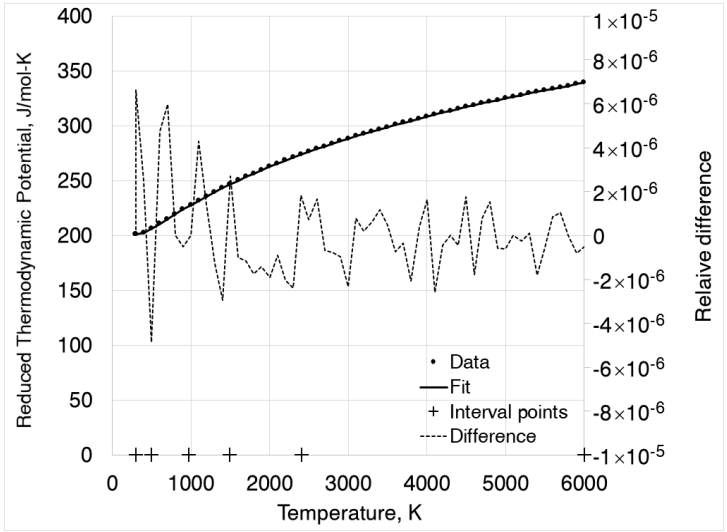
Ethyne (C_2H_2) – C2H2(g)

MW = 26.038 g/mol – $\Delta_f H^o$(298.15 K) = 226.731 kJ/mol – S^o(298.15 K) = 200.958 J/mol-K – p^o = 0.1 MPa



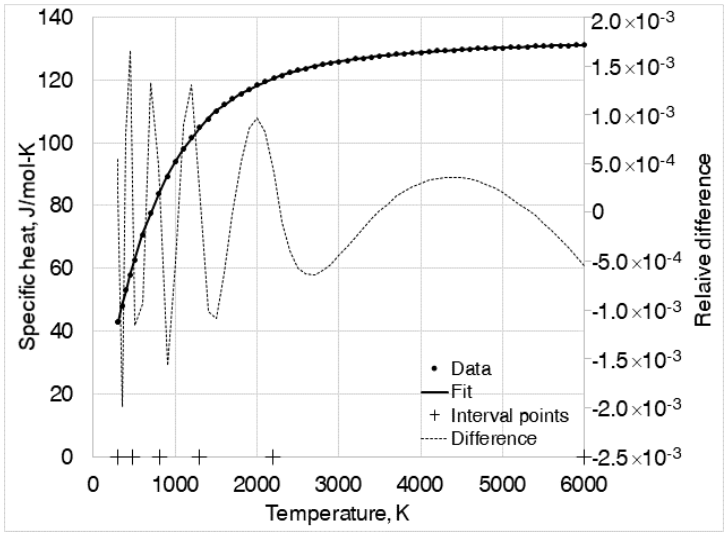


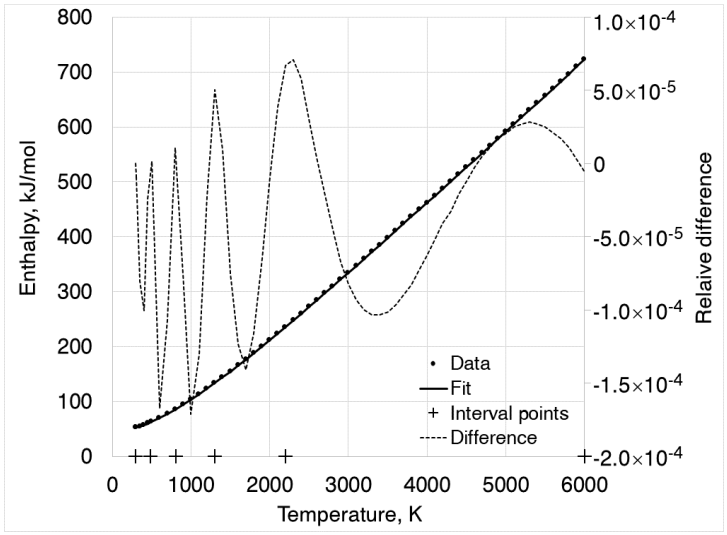


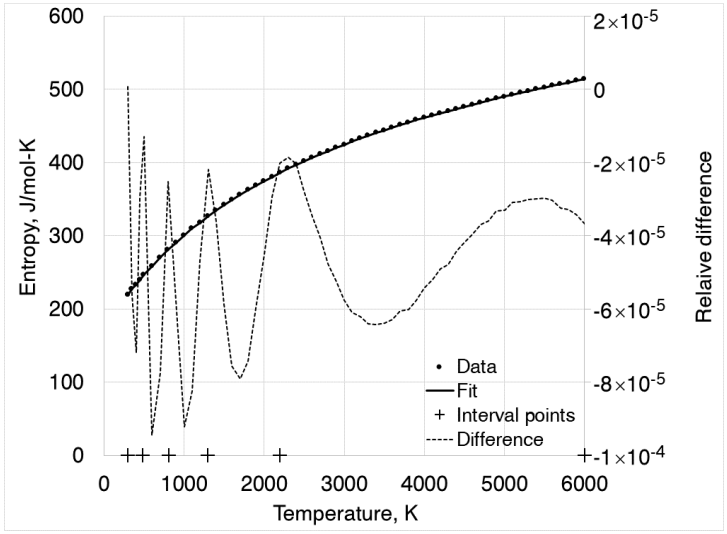


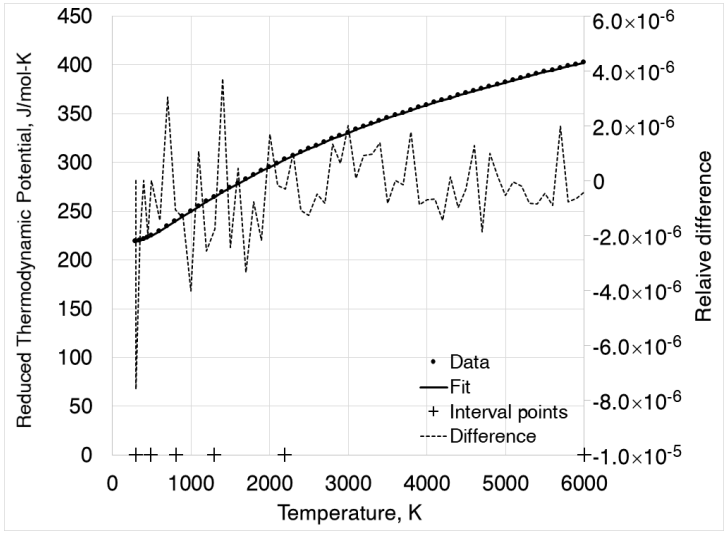
Ethene (C_2H_4) – C2H4(g)

MW = 28.054 g/mol – $\Delta_f H^o$(298.15 K) = 52.467 kJ/mol – S^o(298.15 K) = 219.33 J/mol-K – p^o = 0.1 MPa



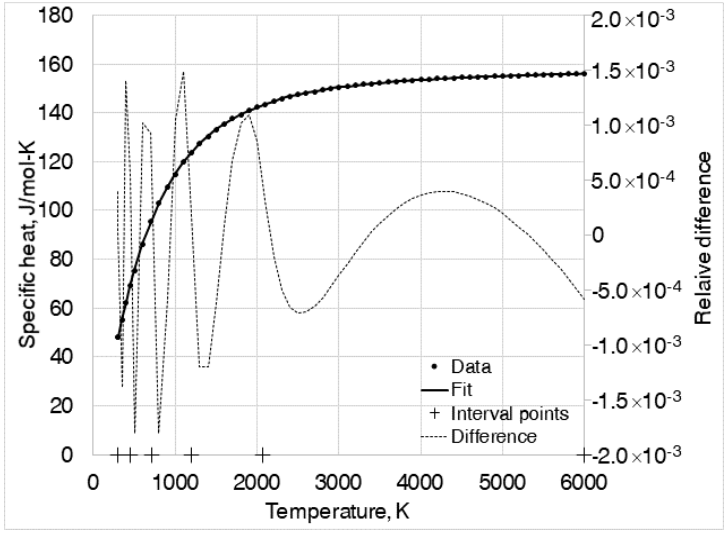


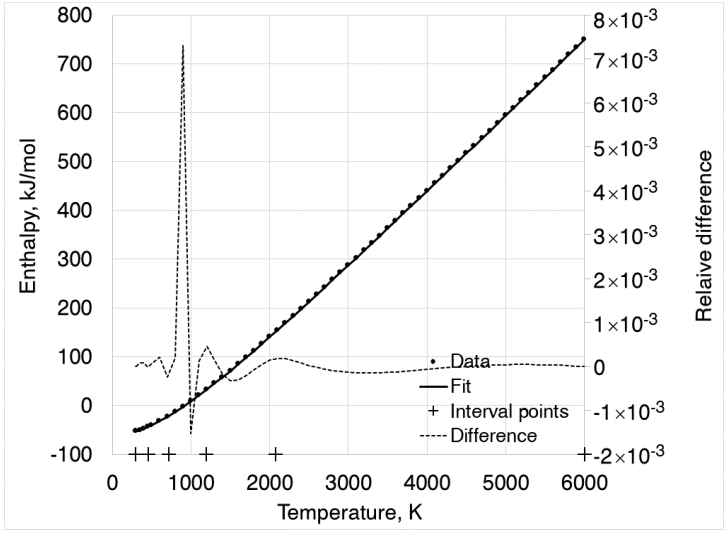


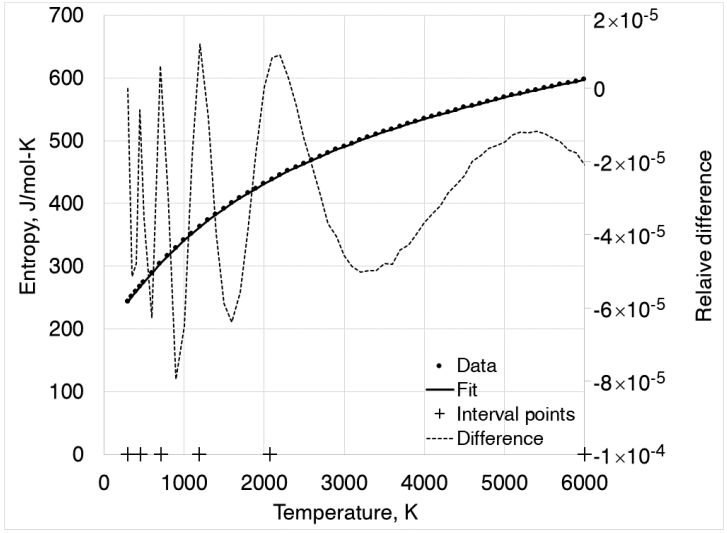


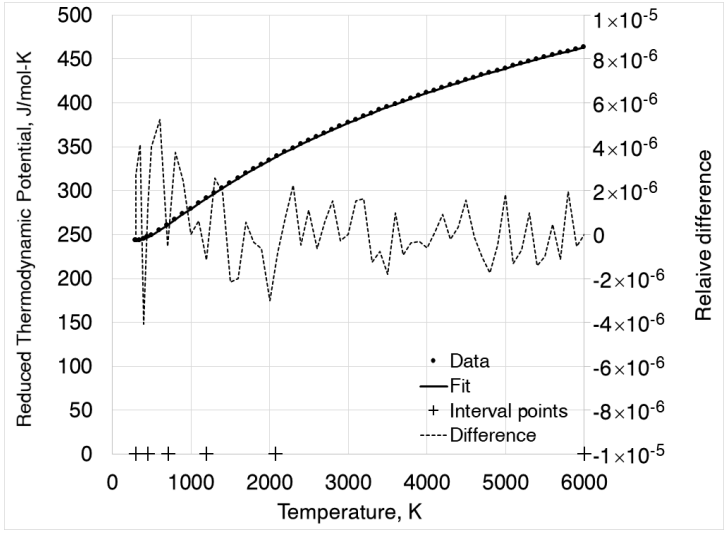
Oxirane (C_2H_4O) – C2H4O1(g)

MW = 44.053 g/mol – $\Delta_f H^o$(298.15 K) = -52.635 kJ/mol – S^o(298.15 K) = 243.012 J/mol-K – p^o = 0.1 MPa



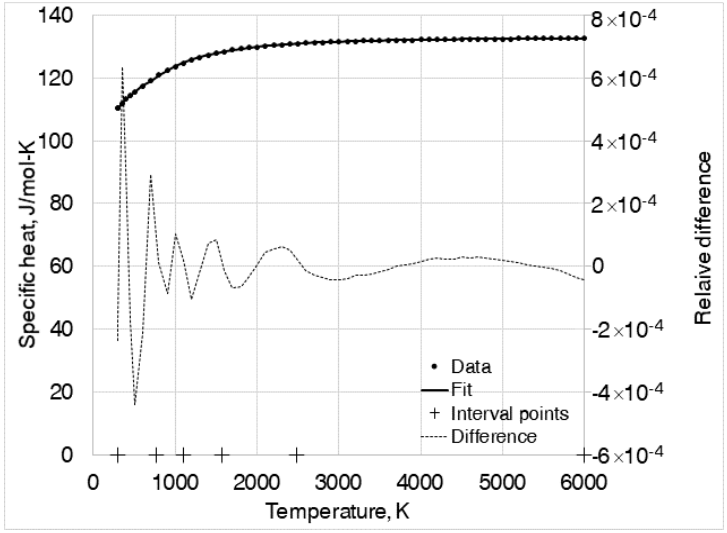


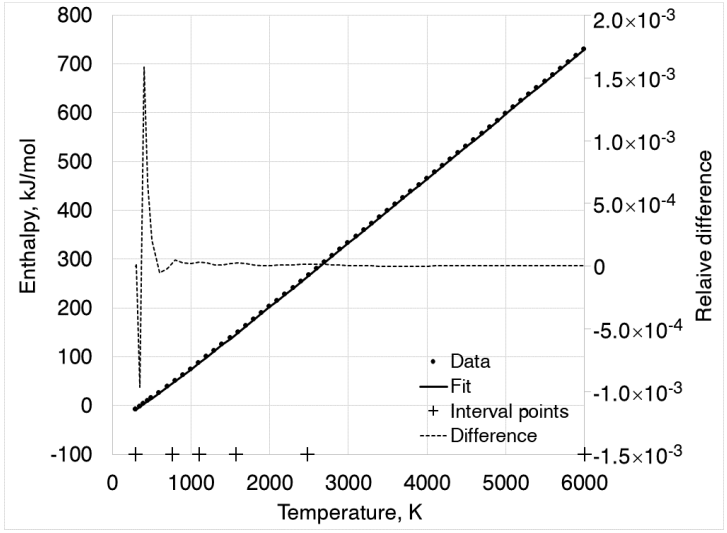


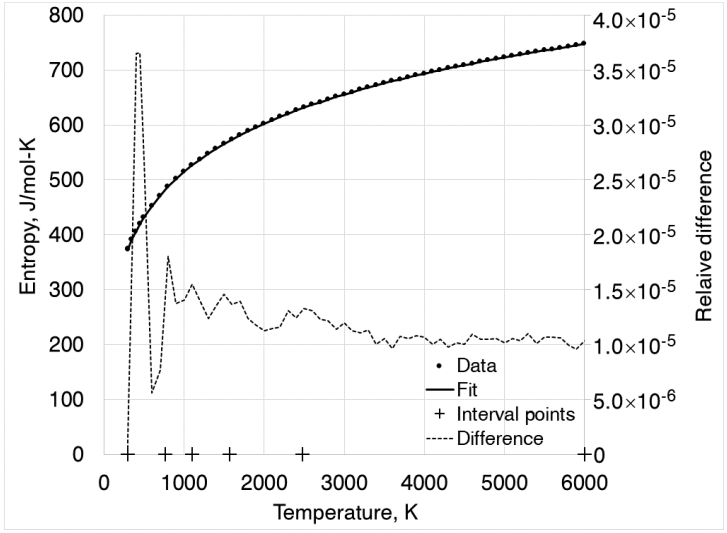


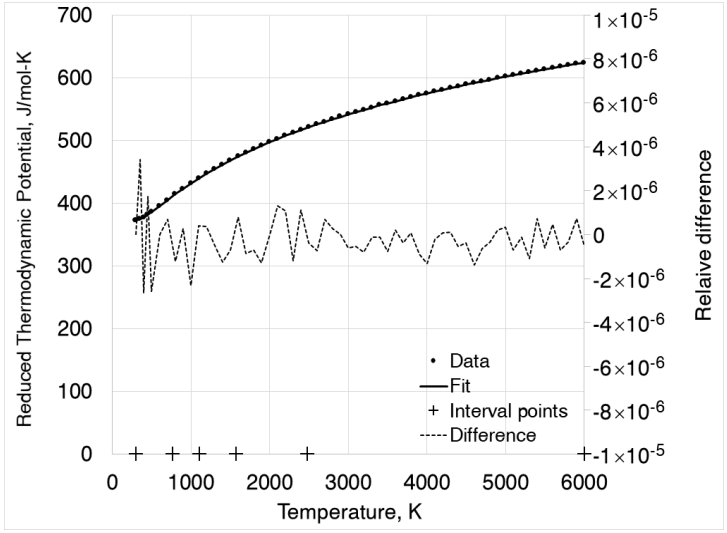
Potassium Cyanide ($(KCN)_2$) – C2K2N2(g)

MW = 130.232 g/mol – $\Delta_f H^o$(298.15 K) = -8.368 kJ/mol – S^o(298.15 K) = 373.129 J/mol-K – p^o = 0.1 MPa



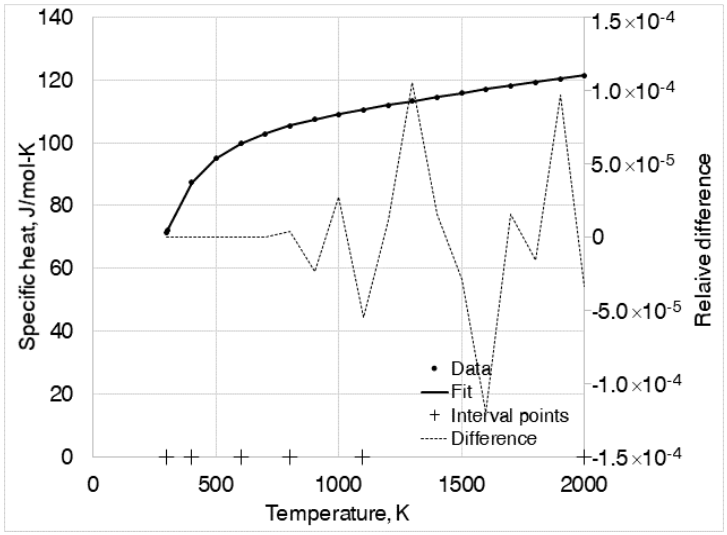


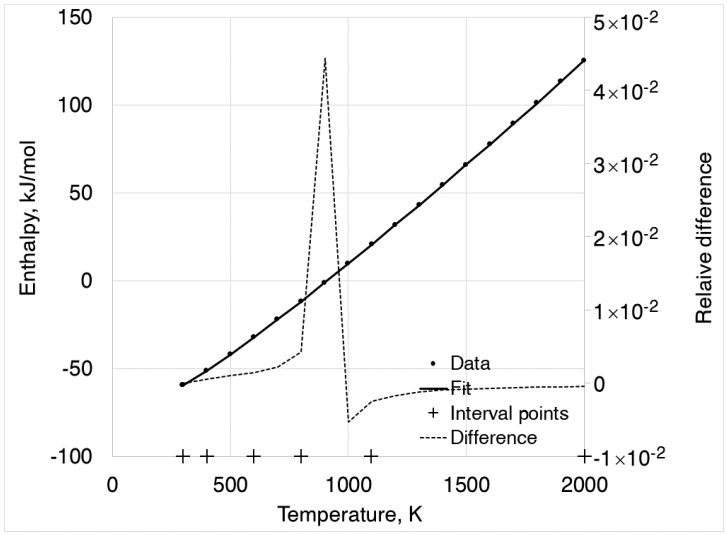


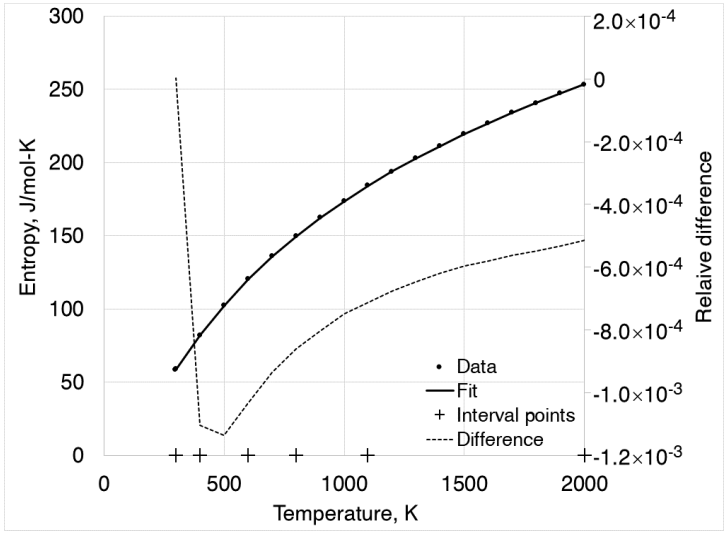


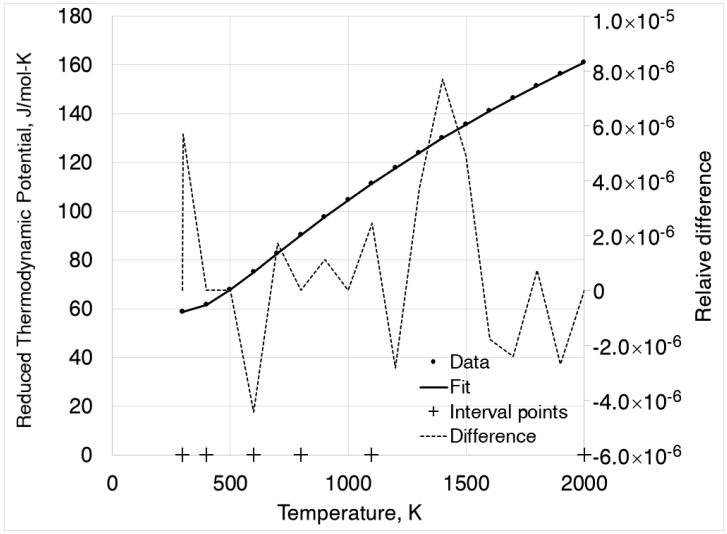
Lithium Carbide (Li_2C_2) – C2Li2(cr)

MW = 37.904 g/mol – $\Delta_f H^o$(298.15 K) = -59.413 kJ/mol – S^o(298.15 K) = 58.576 J/mol-K – p^o = 0.1 MPa



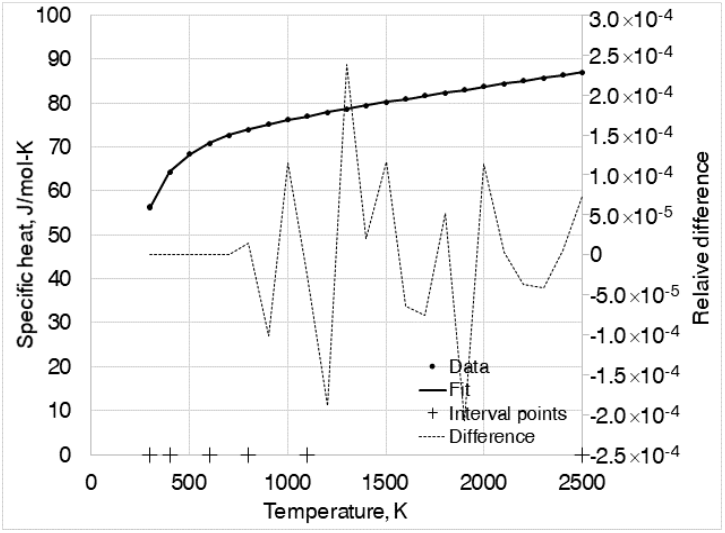


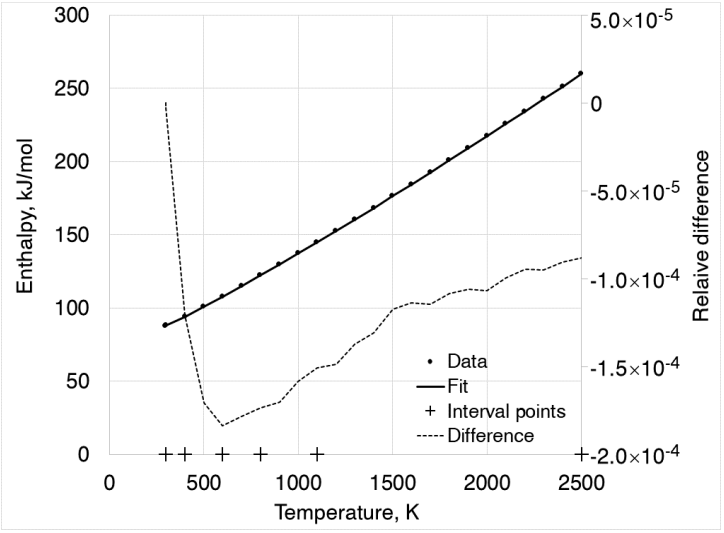


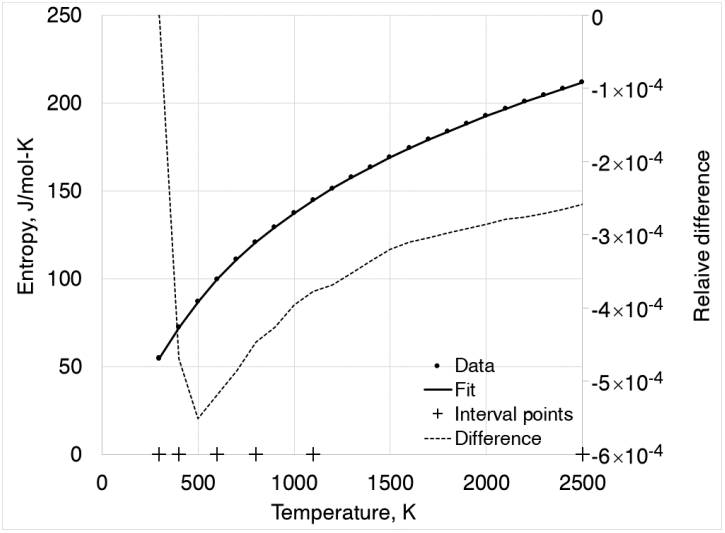


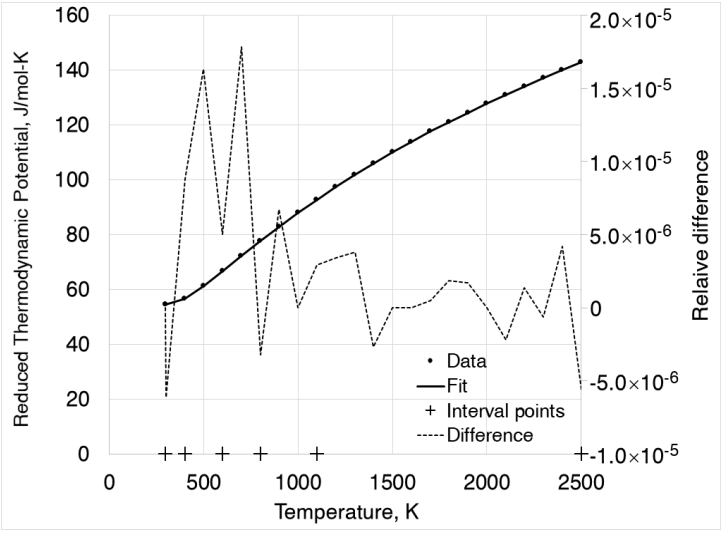
Magnesium Carbide (MgC_2) – C2Mg1(cr)

MW = 48.327 g/mol – $\Delta_f H^o$(298.15 K) = 87.864 kJ/mol – S^o(298.15 K) = 54.392 J/mol-K – p^o = 0.1 MPa



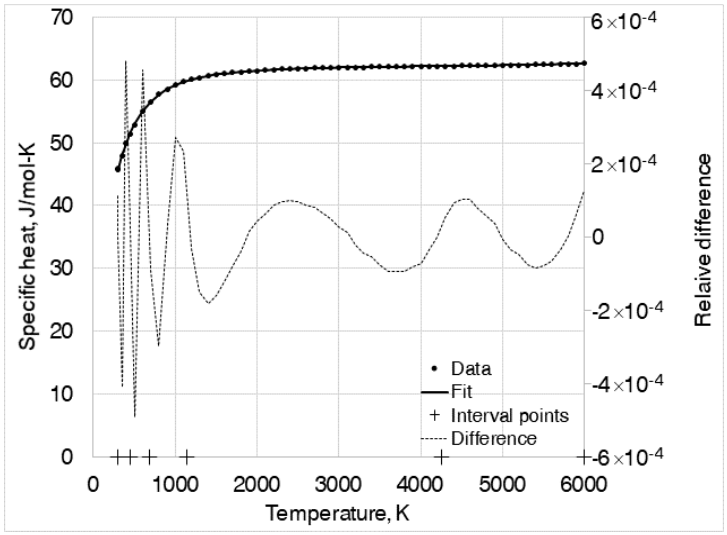


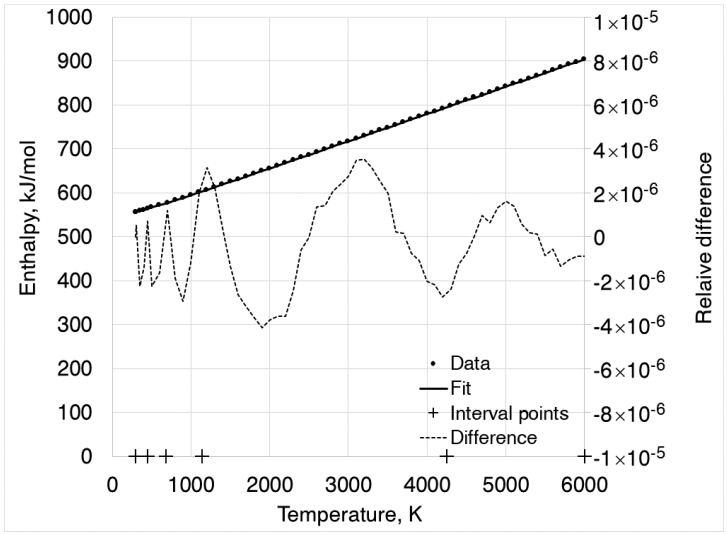


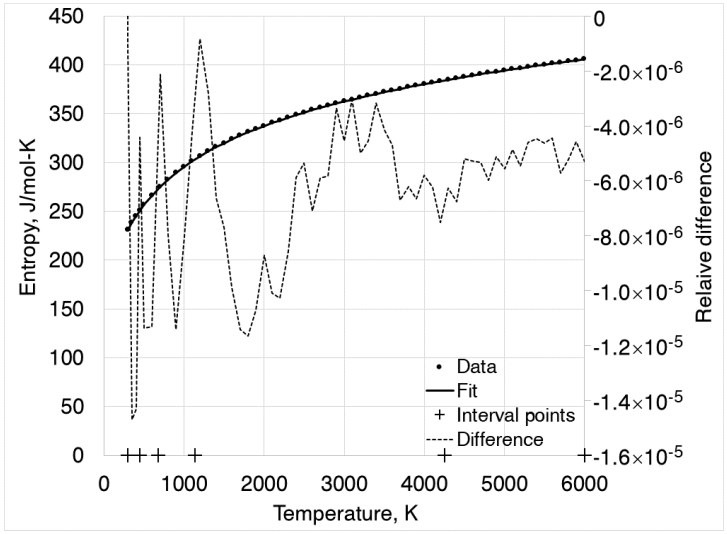


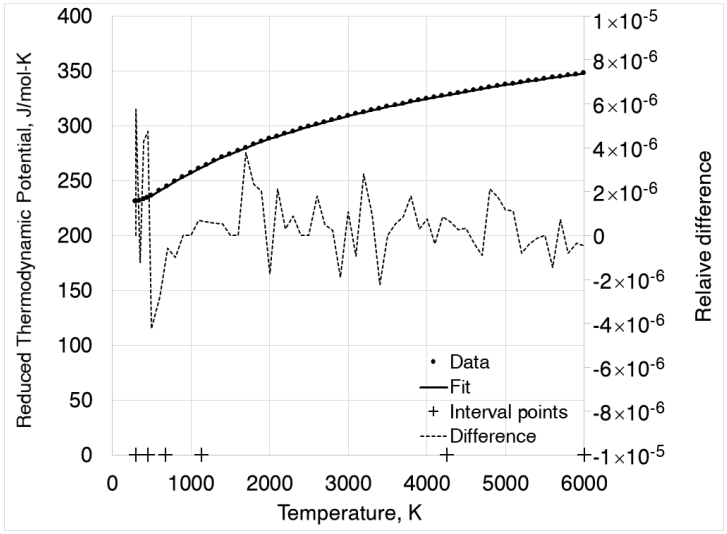
CNC Radical (C_2N) – C2N1(g)

MW = 38.029 g/mol – $\Delta_f H^o$(298.15 K) = 556.472 kJ/mol – S^o(298.15 K) = 230.905 J/mol-K – p^o = 0.1 MPa



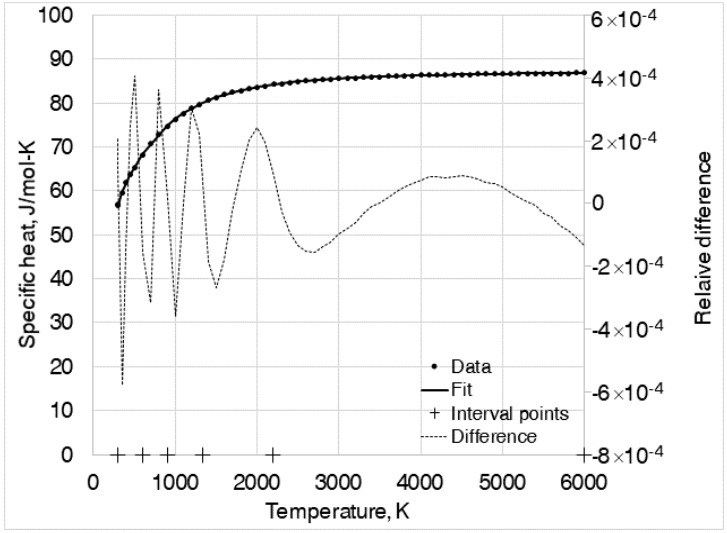


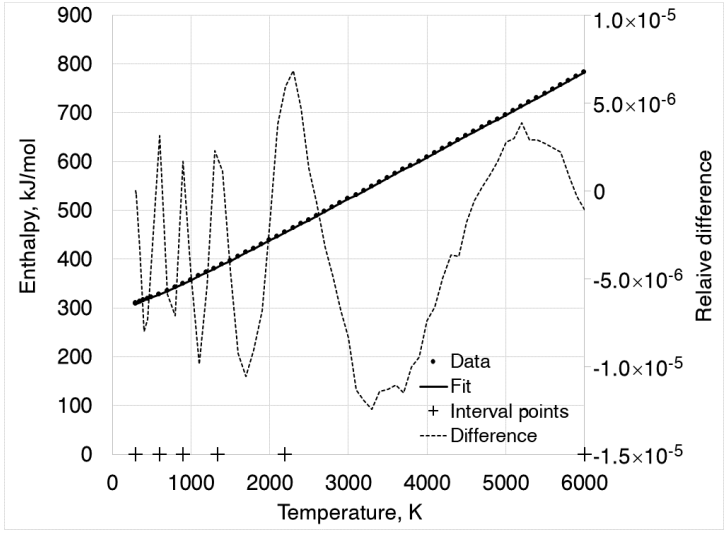


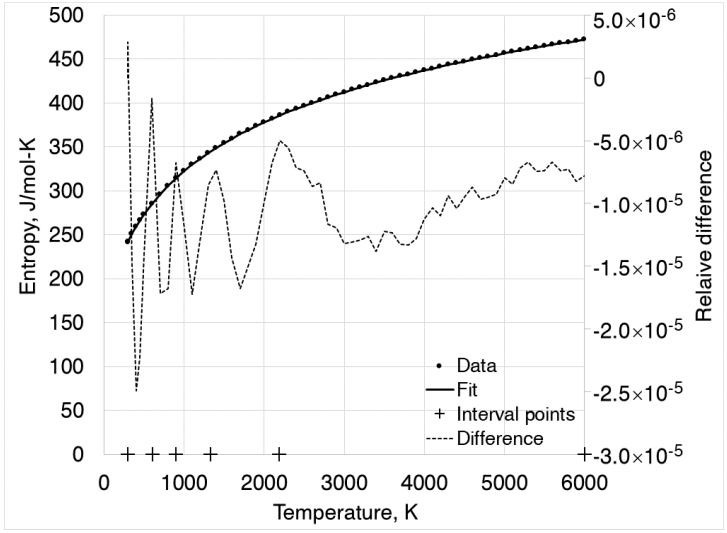


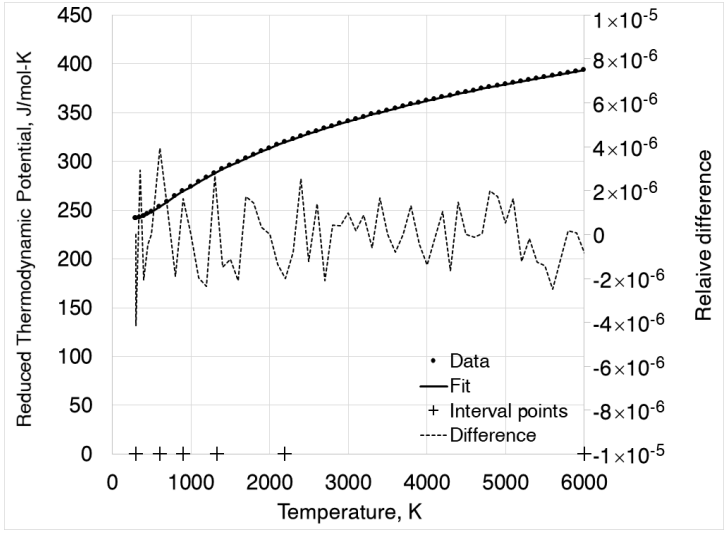
Ethanedinitrile ($(CN)_2$) – C2N2(g)

MW = 52.035 g/mol – $\Delta_f H^o$(298.15 K) = 309.072 kJ/mol – S^o(298.15 K) = 241.565 J/mol-K – p^o = 0.1 MPa



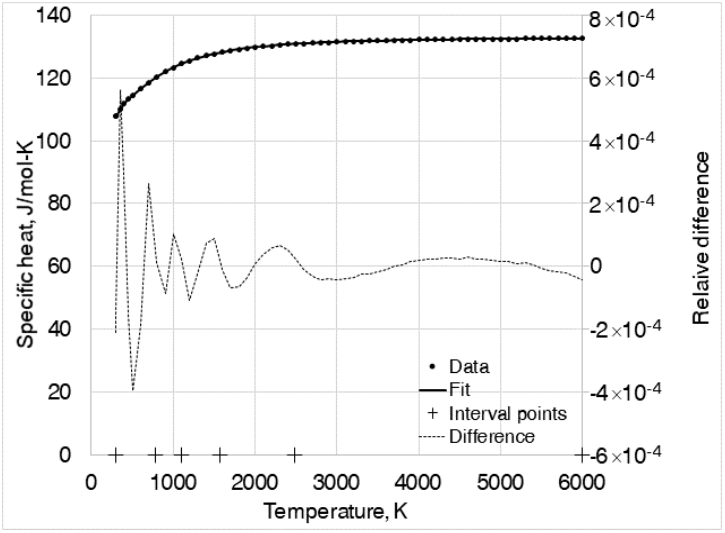


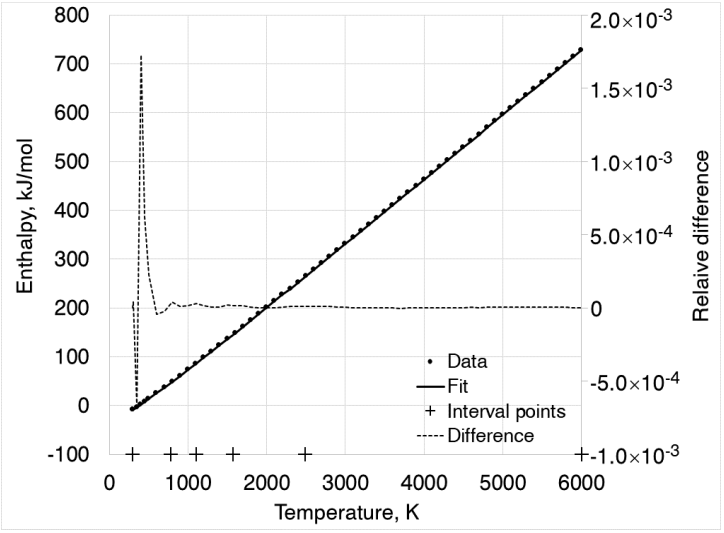


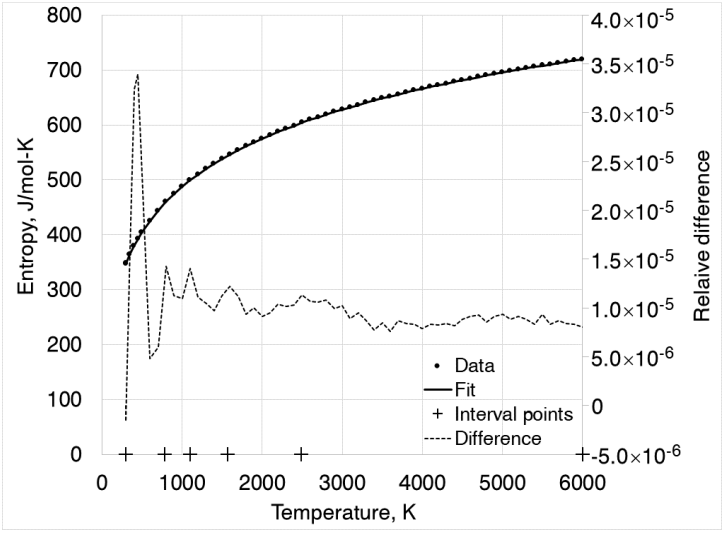


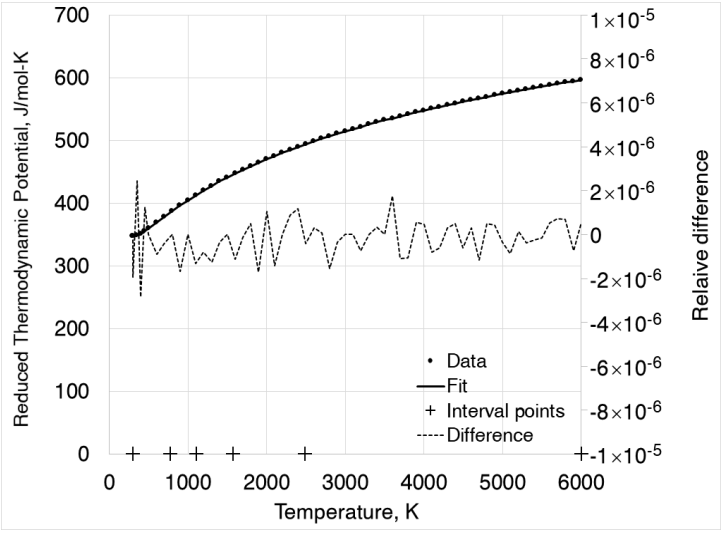
Sodium Cyanide ((NaCN)$_2$) – C2N2Na2(g)

MW = 98.015 g/mol – $\Delta_f H^o$(298.15 K) = -8.786 kJ/mol – S^o(298.15 K) = 347.102 J/mol-K – p^o = 0.1 MPa



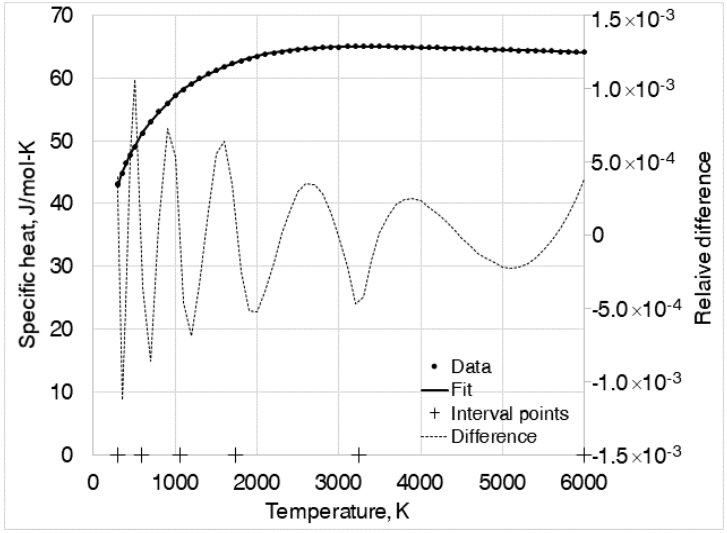


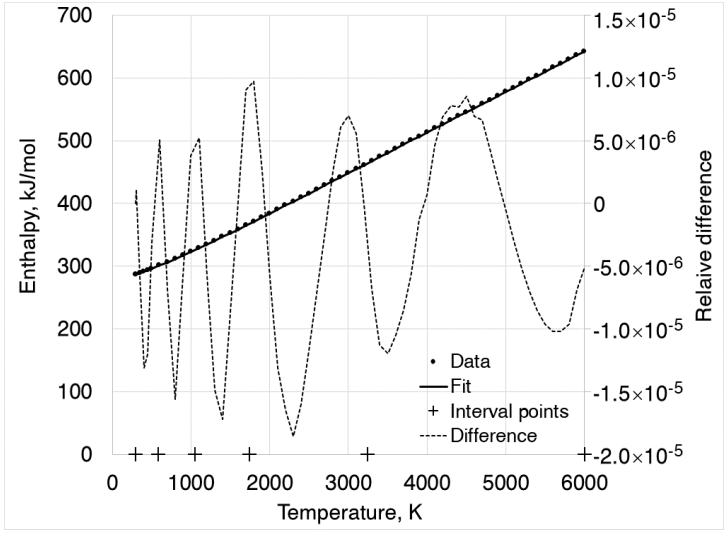


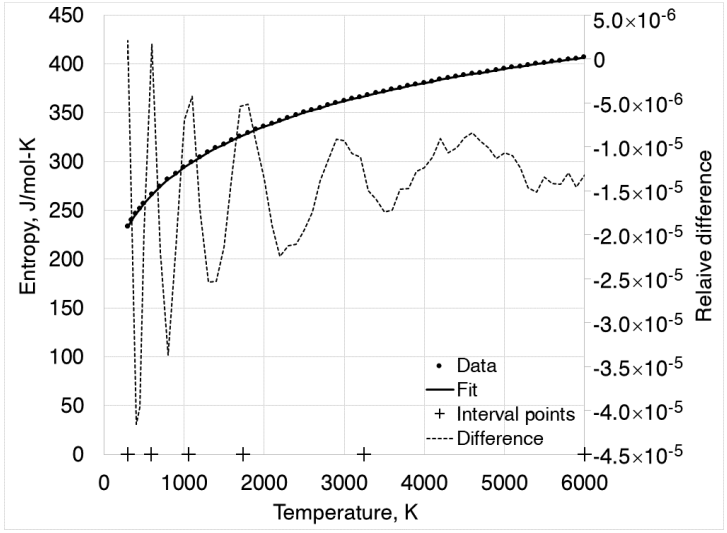


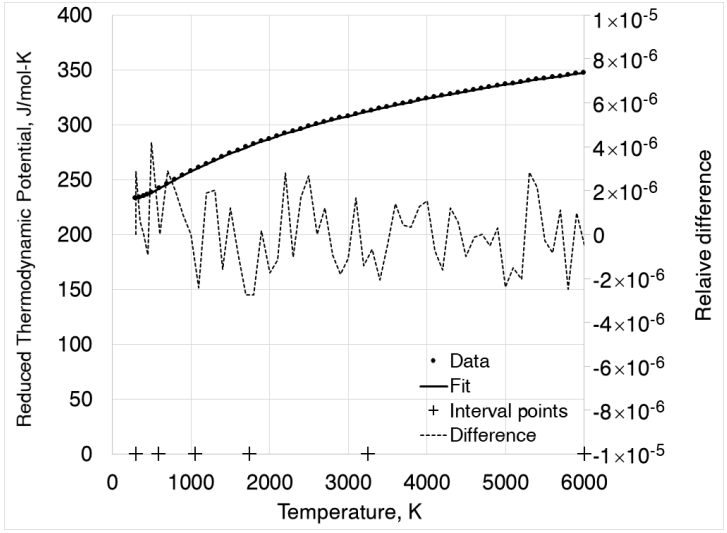
CCO Radical (C_2O) – C2O1(g)

MW = 40.021 g/mol – $\Delta_f H^o$(298.15 K) = 286.604 kJ/mol – S^o(298.15 K) = 233.066 J/mol-K – p^o = 0.1 MPa



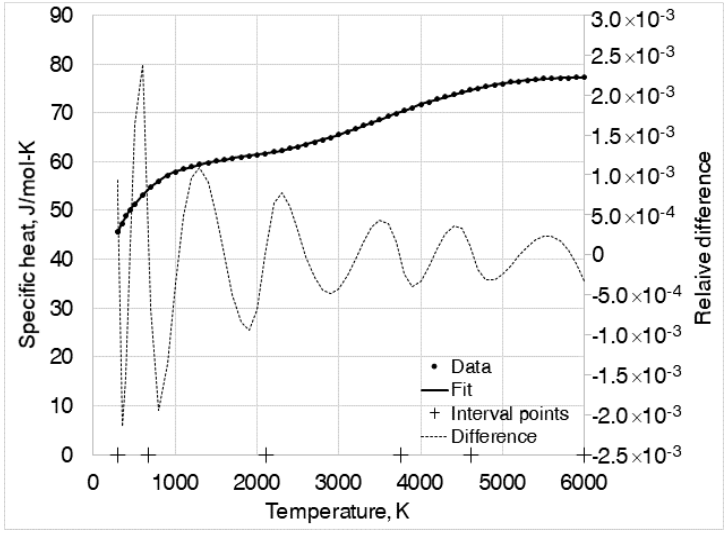


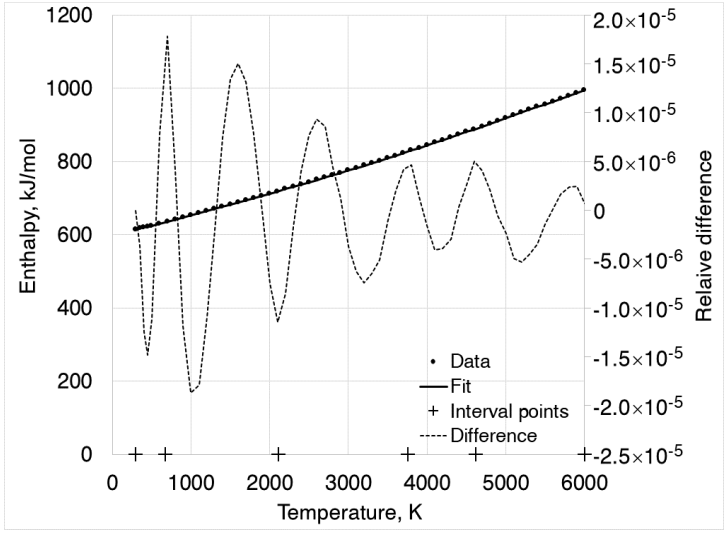


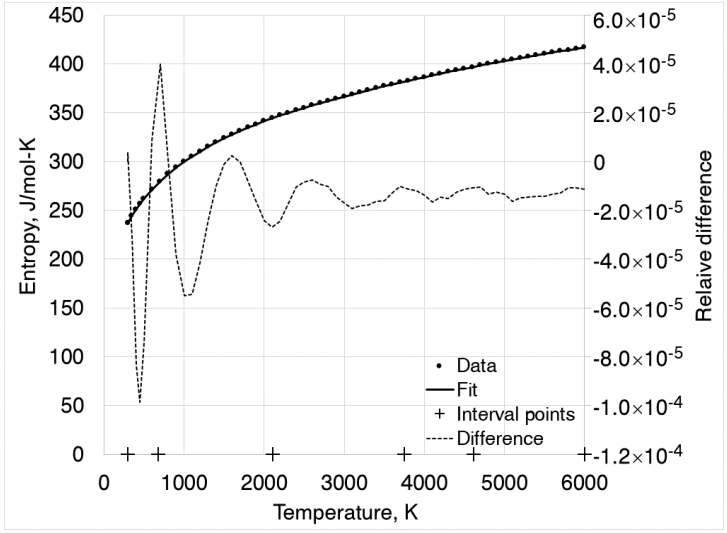


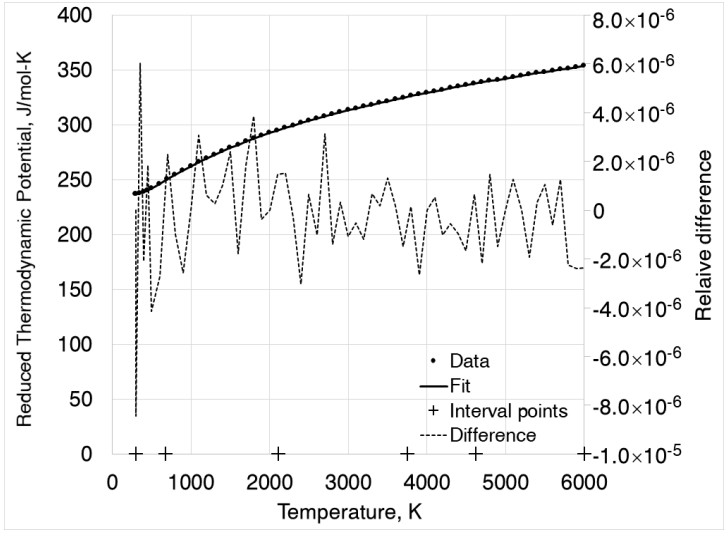
Silicon Carbide (SiC_2) – C2Si1(g)

MW = 52.108 g/mol – $\Delta_f H^o$(298.15 K) = 615.048 kJ/mol – S^o(298.15 K) = 236.731 J/mol-K – p^o = 0.1 MPa



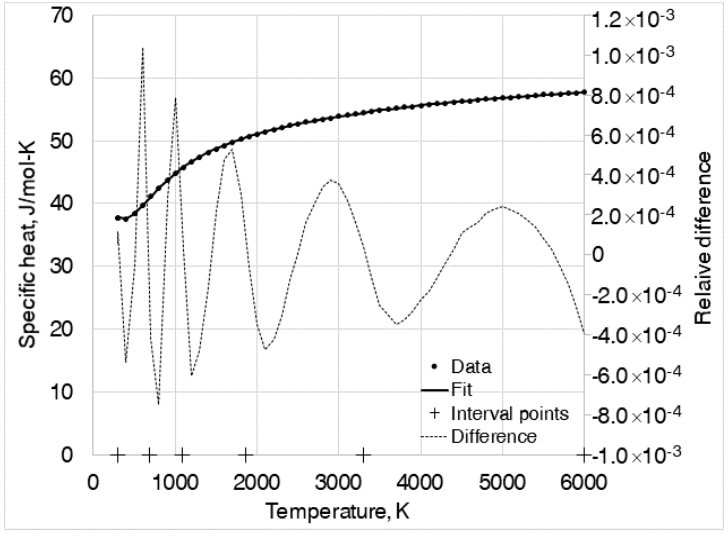


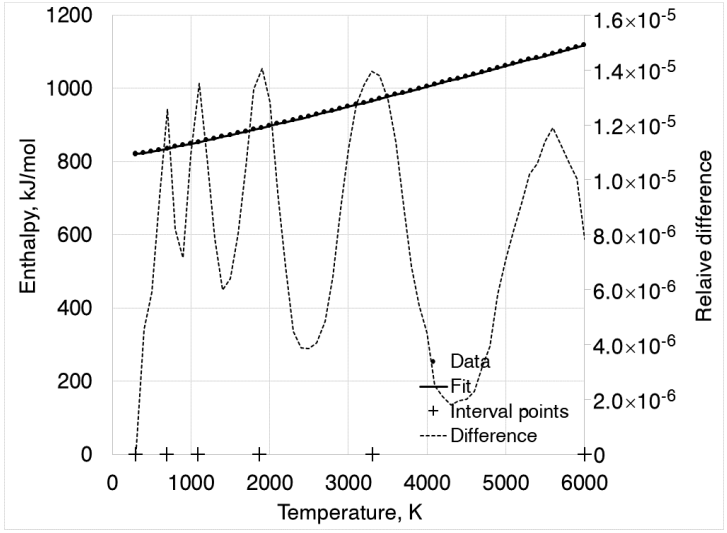


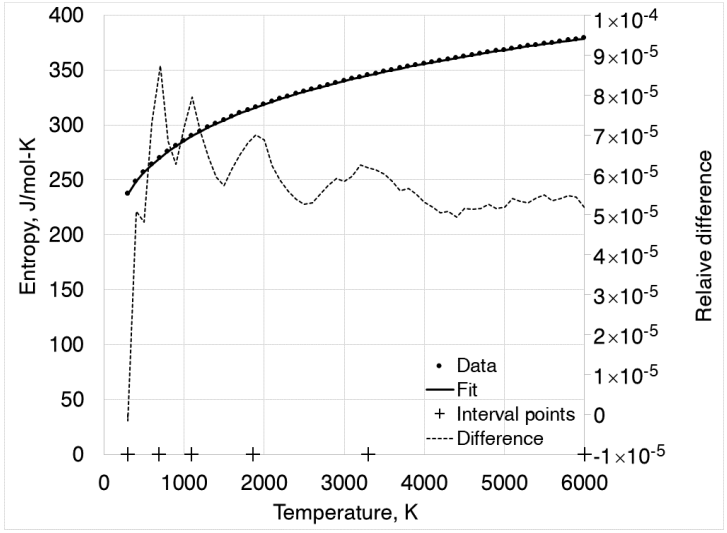


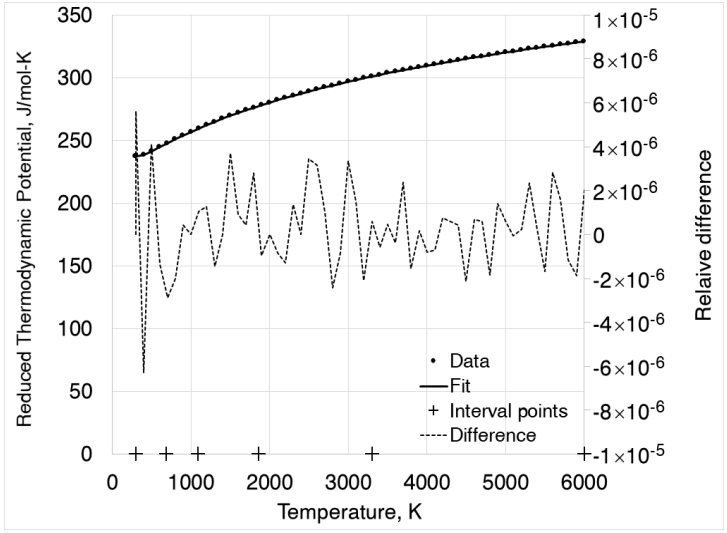
Carbon (C_3) – C3(g)

MW = 36.033 g/mol – $\Delta_f H^o$(298.15 K) = 820.064 kJ/mol – S^o(298.15 K) = 237.246 J/mol-K – p^o = 0.1 MPa



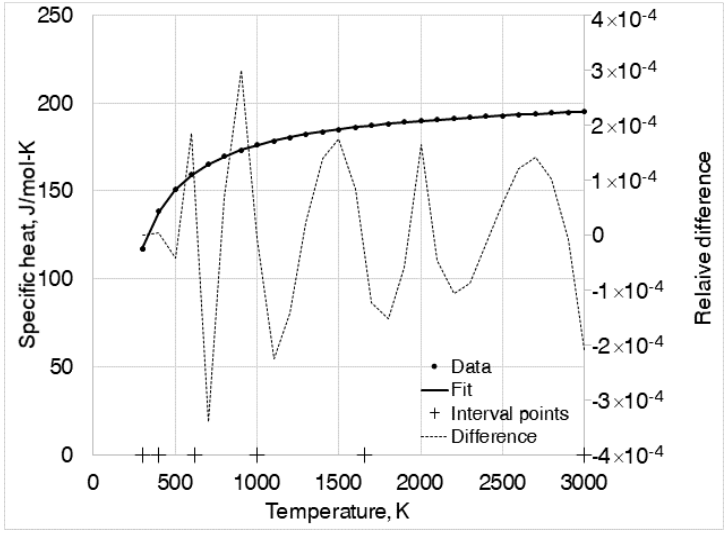


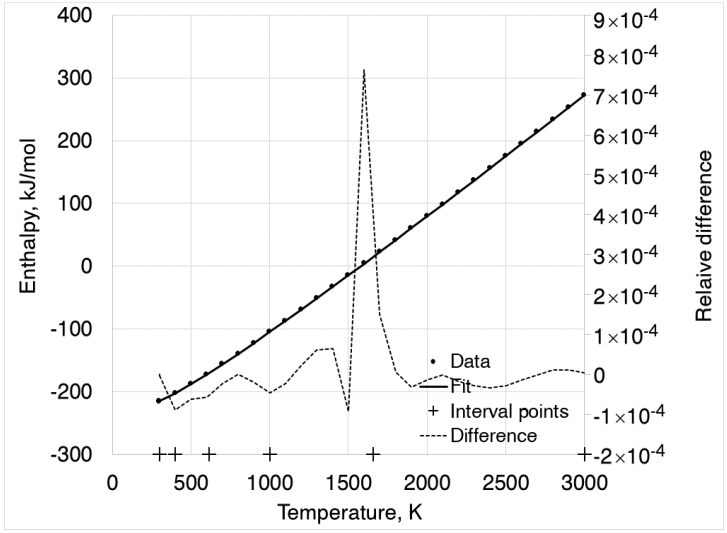


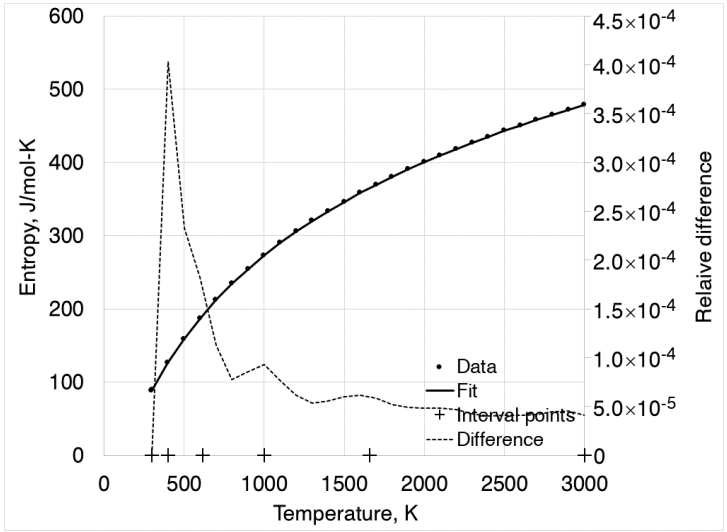


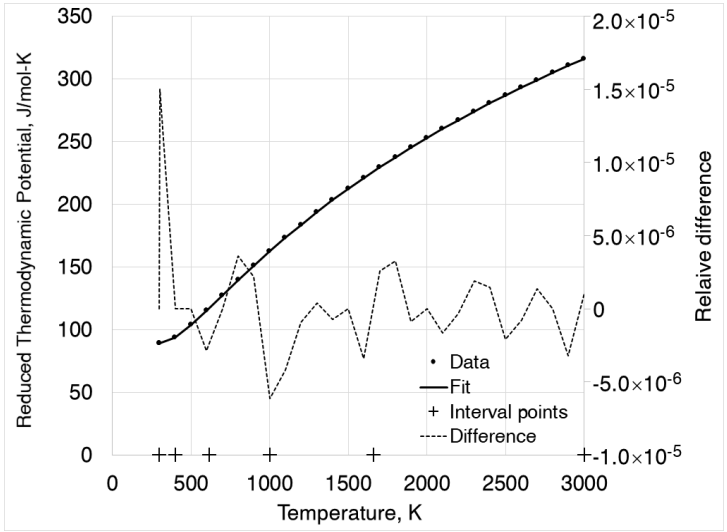
Aluminum Carbide (Al_4C_3) – C3Al4(cr)

MW = 143.959 g/mol – $\Delta_f H^o$(298.15 K) = -215.685 kJ/mol – S^o(298.15 K) = 88.969 J/mol-K – p^o = 0.1 MPa



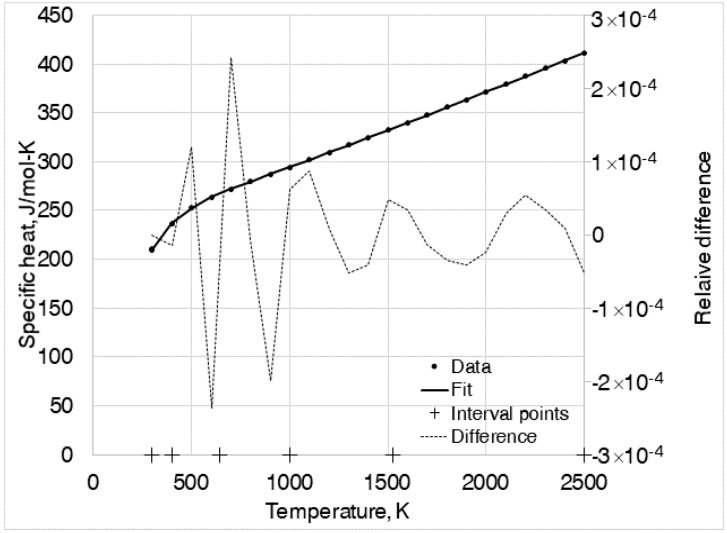


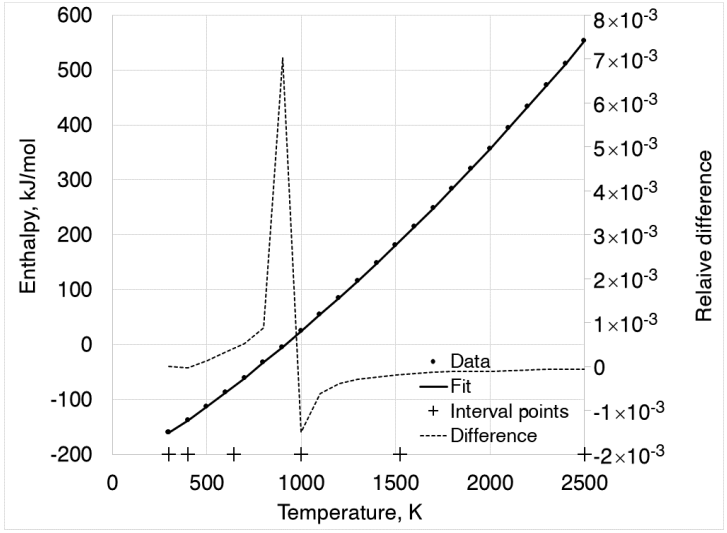


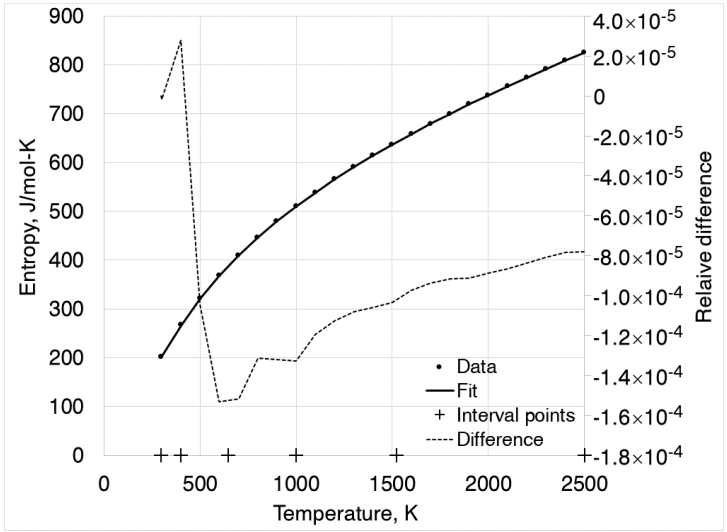


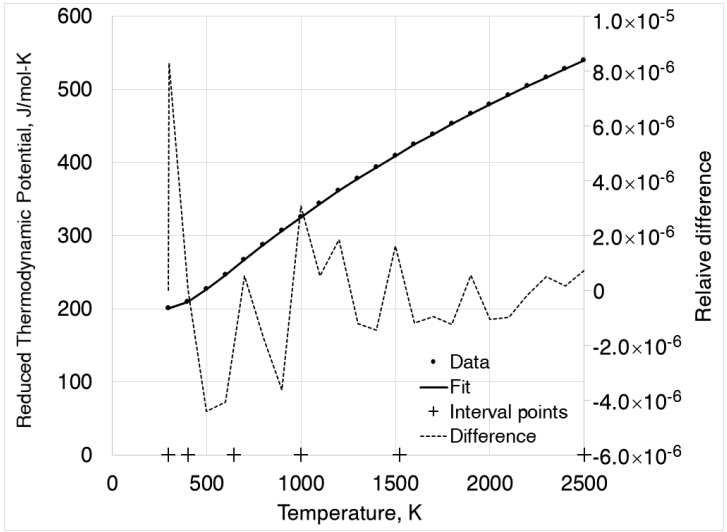
Chromium Carbide (Cr_7C_3) – C3Cr7(cr)

MW = 400.005 g/mol – $\Delta_f H^o$(298.15 K) = -160.666 kJ/mol – S^o(298.15 K) = 200.983 J/mol-K – p^o = 0.1 MPa



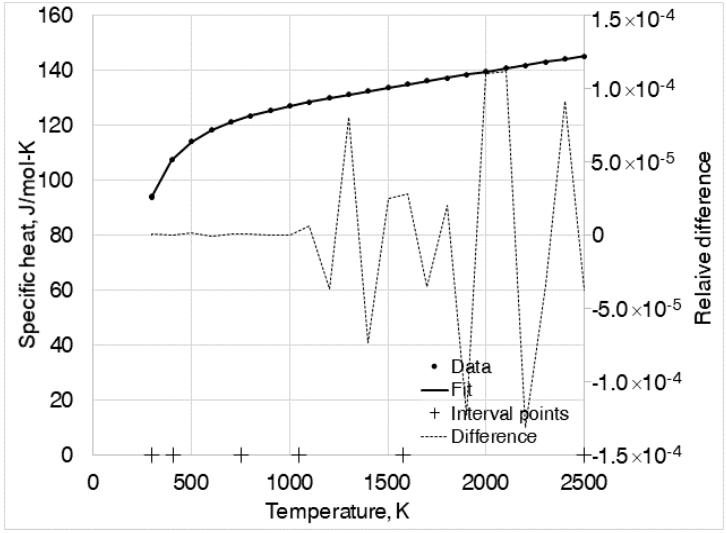


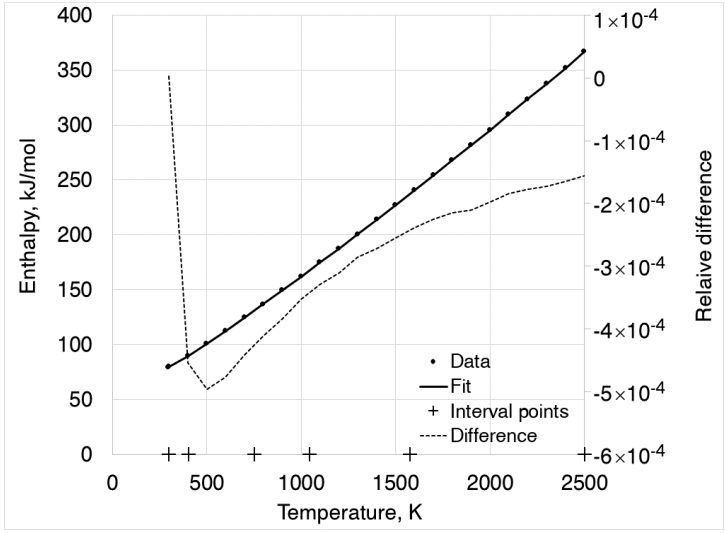


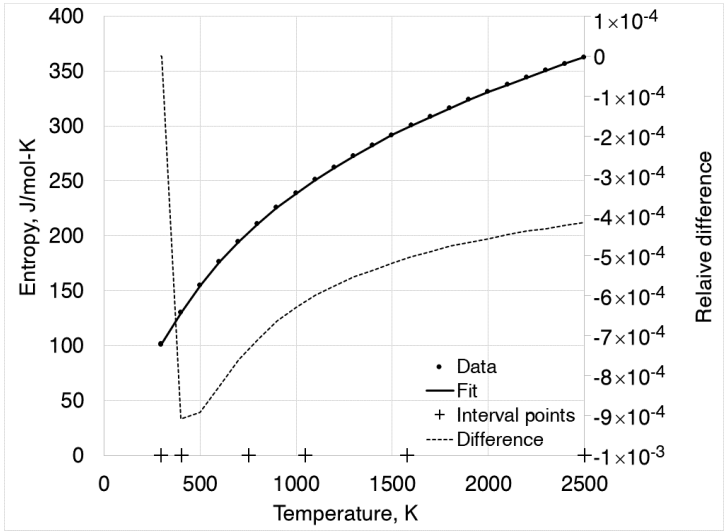


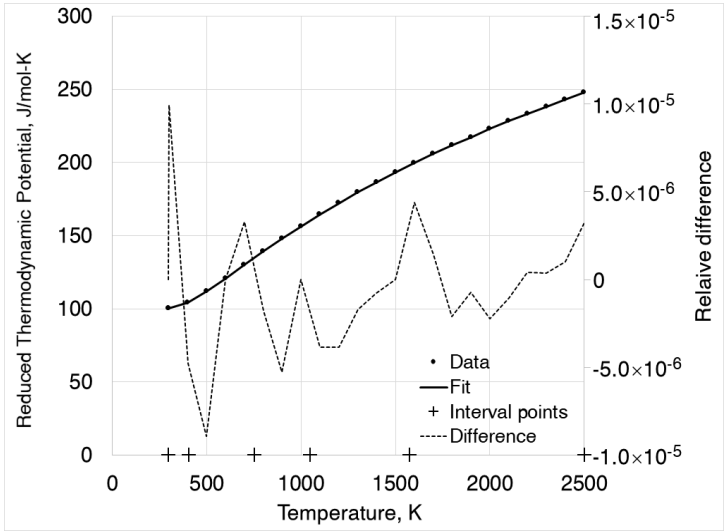
Magnesium Carbide (Mg_2C_3) – C3Mg2(cr)

MW = 84.643 g/mol – $\Delta_f H^o$(298.15 K) = 79.496 kJ/mol – S^o(298.15 K) = 100.416 J/mol-K – p^o = 0.1 MPa



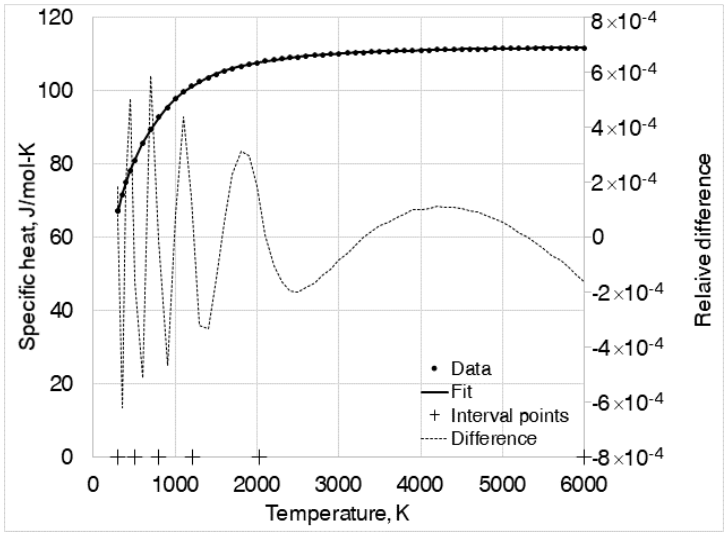


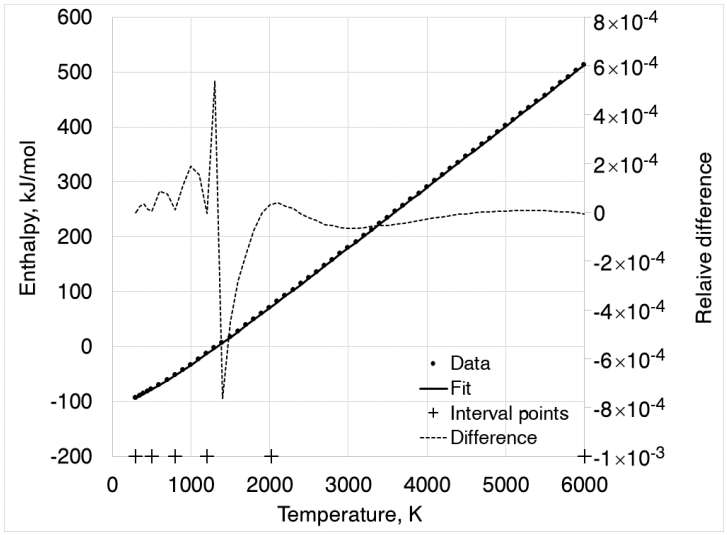


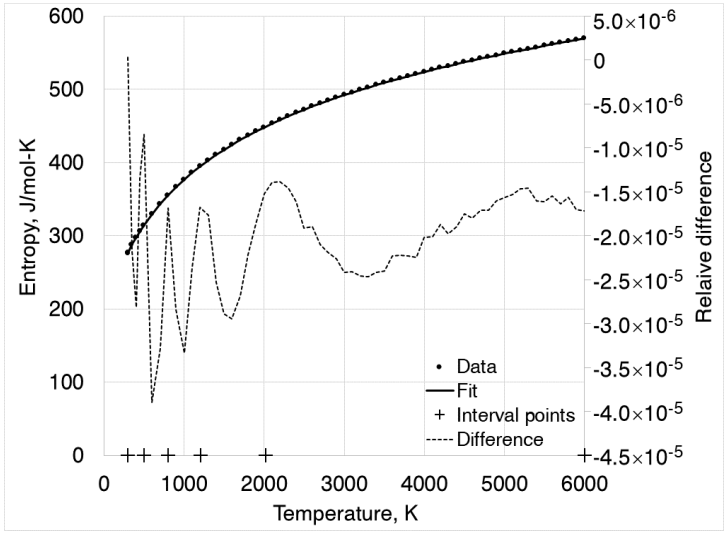


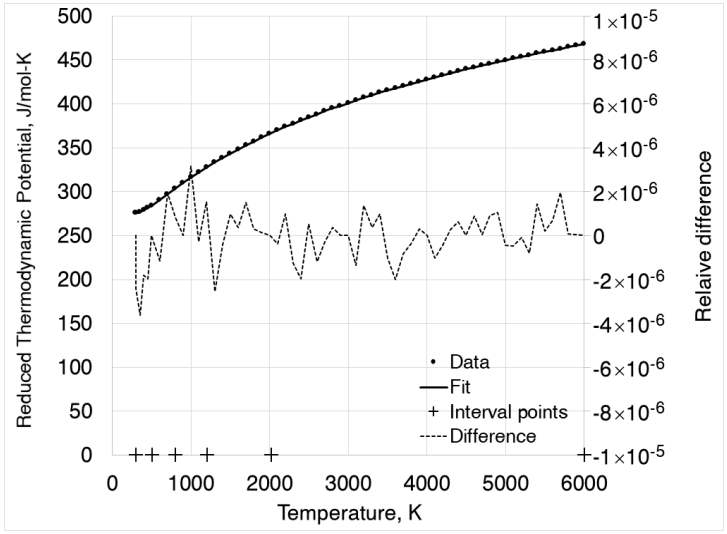
Carbon Suboxide (C_3O_2) – C3O2(g)

MW = 68.032 g/mol – $\Delta_f H^o$(298.15 K) = -93.638 kJ/mol – S^o(298.15 K) = 276.071 J/mol-K – p^o = 0.1 MPa



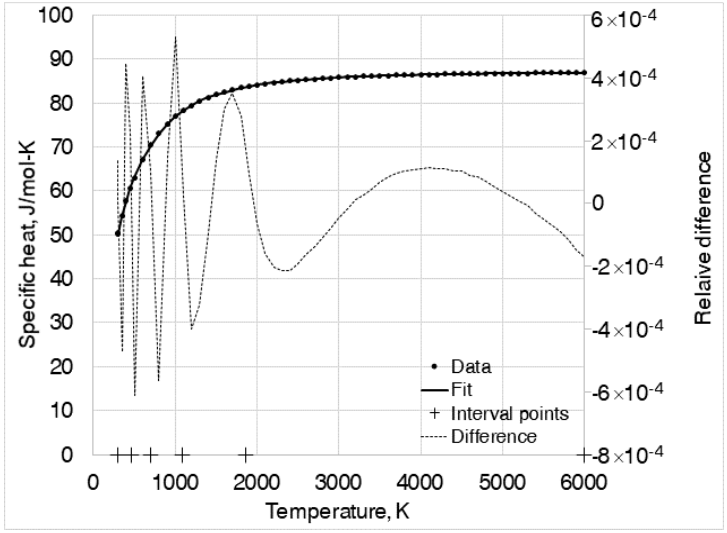


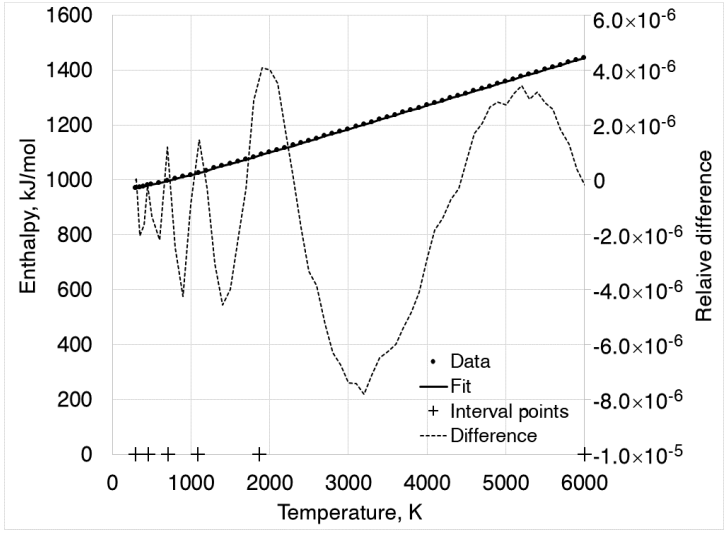


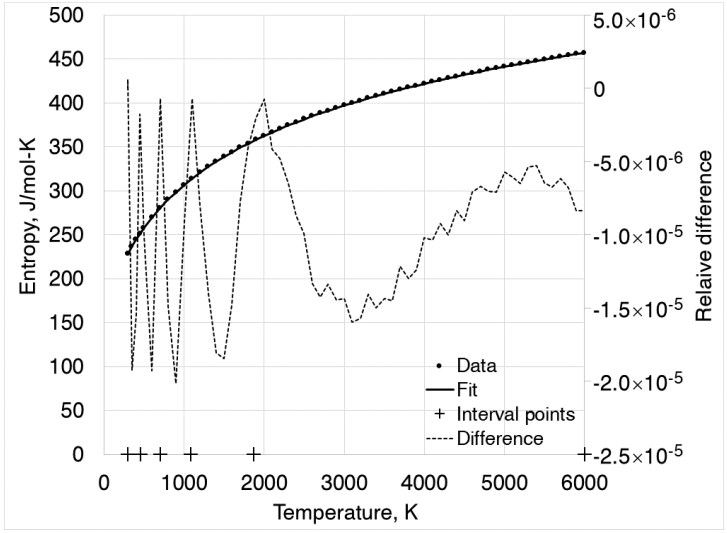


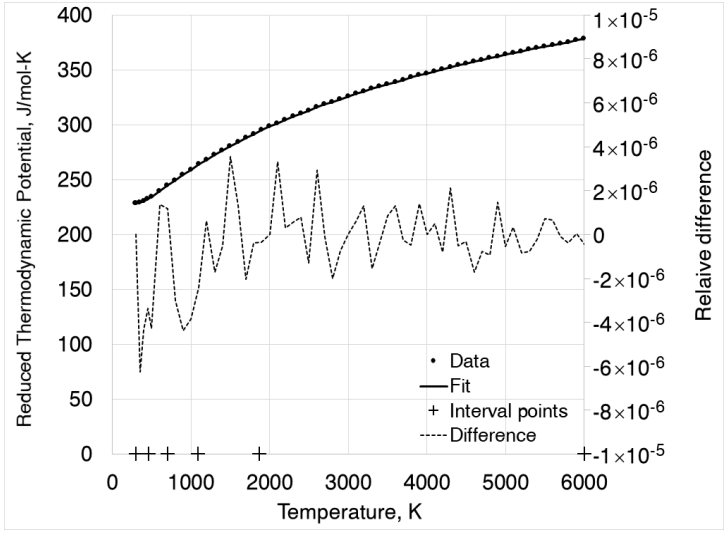
Carbon (C_4) – C4(g)

MW = 48.044 g/mol – $\Delta_f H^o$(298.15 K) = 970.688 kJ/mol – S^o(298.15 K) = 228.322 J/mol-K – p^o = 0.1 MPa



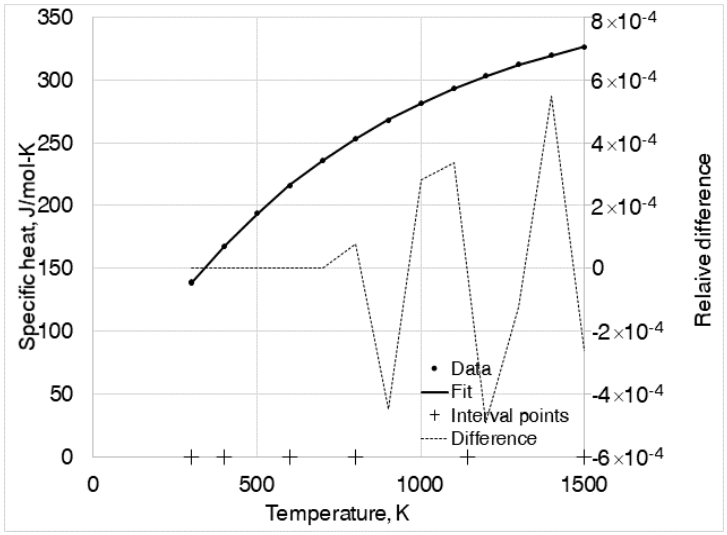


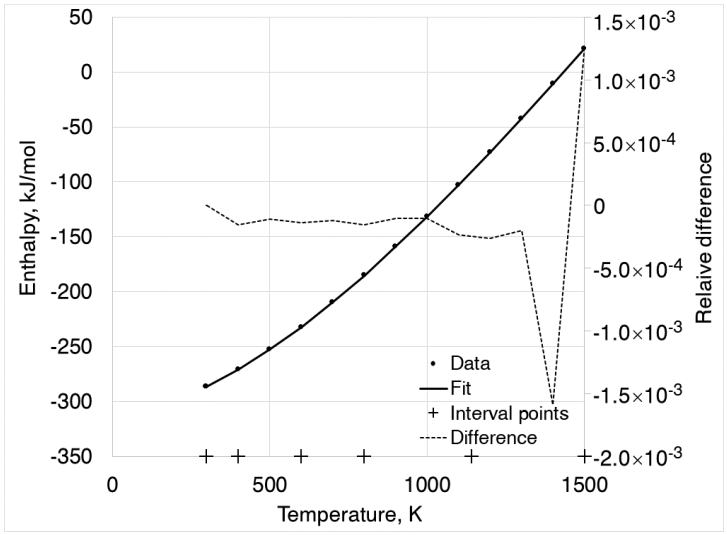


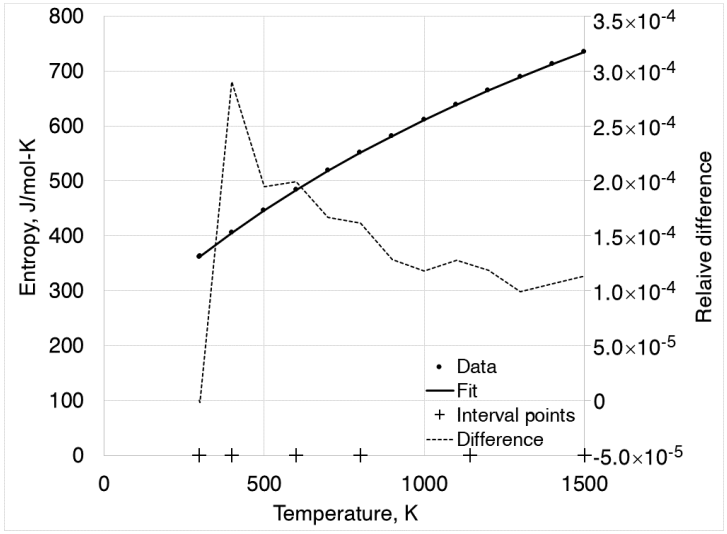


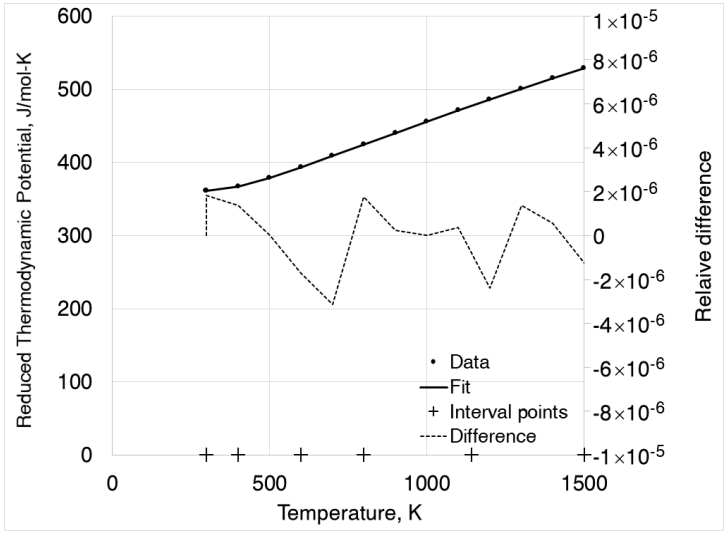
Tetramethylsilane ($Si(CH_3)_4$) – C4H12Si1(g)

MW = 88.225 g/mol – $\Delta_f H^o$(298.15 K) = -286.604 kJ/mol – S^o(298.15 K) = 361.189 J/mol-K – p^o = 0.1 MPa



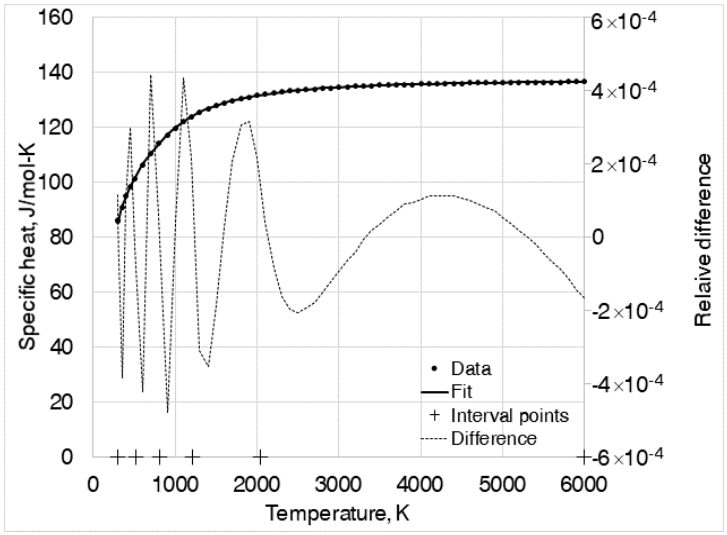


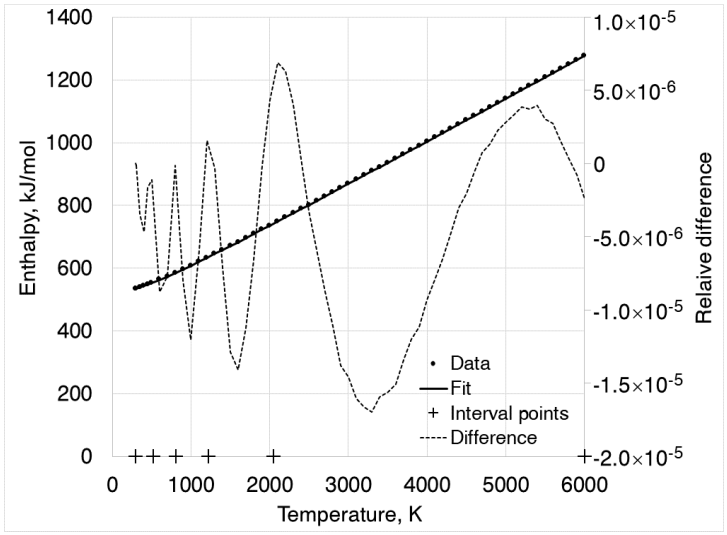


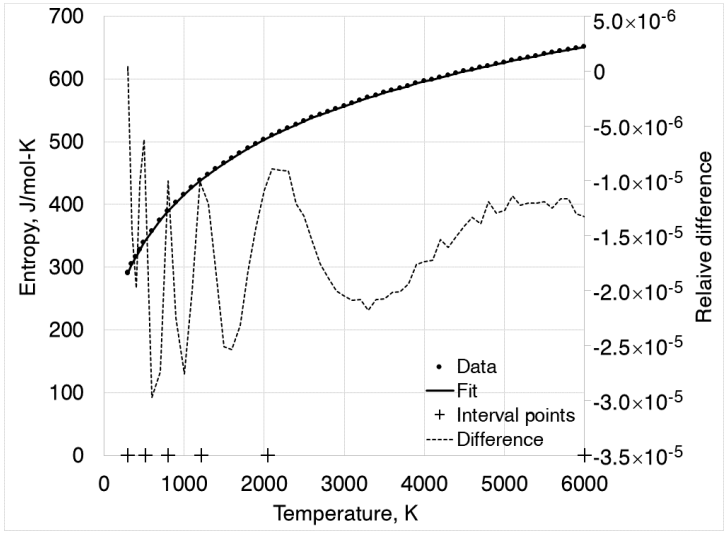


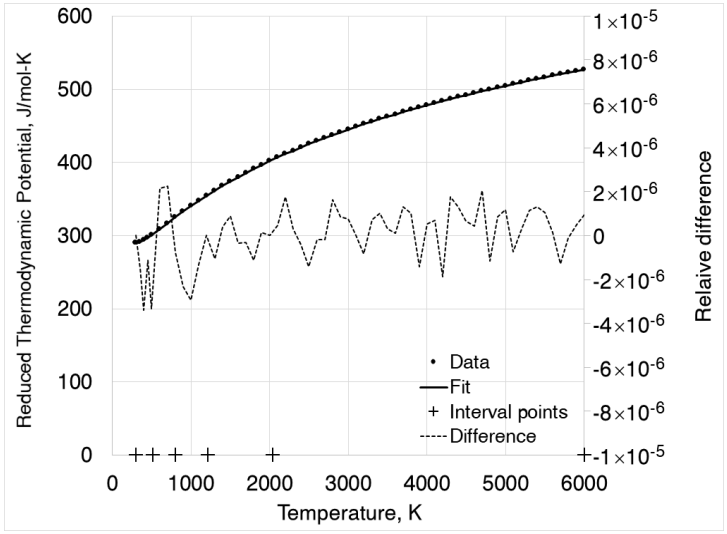
$_2$-Butynedinitrile (C_4N_2) – C4N2(g)

MW = 76.057 g/mol – $\Delta_f H^o$(298.15 K) = 533.46 kJ/mol – S^o(298.15 K) = 290.112 J/mol-K – p^o = 0.1 MPa



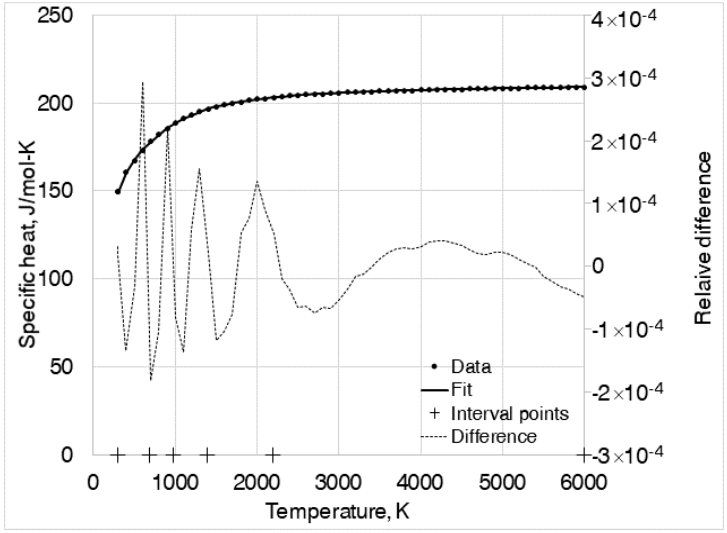


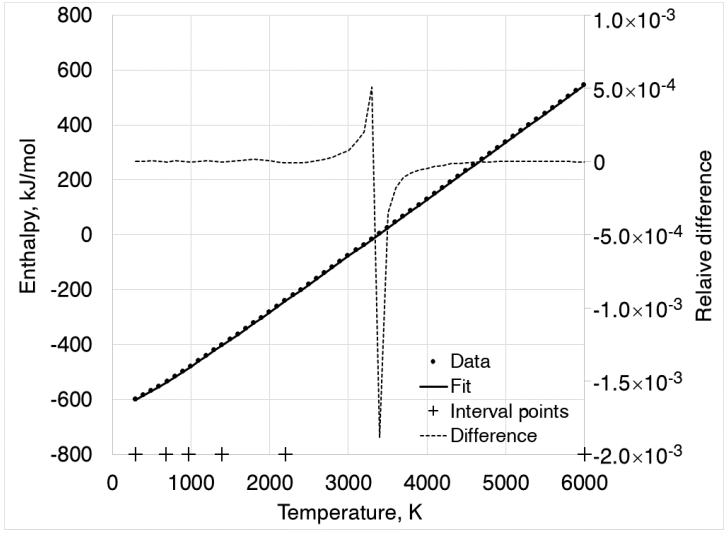


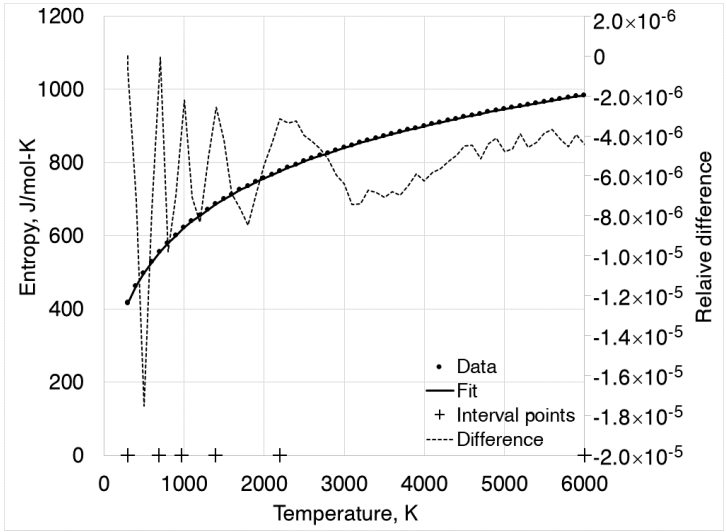


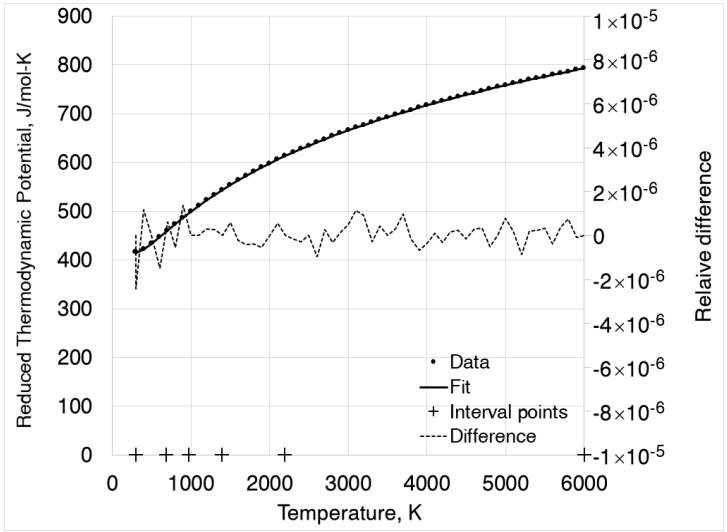
Nickel Carbonyl ($Ni(CO)_4$) – C4Ni1O4(g)

MW = 170.73 g/mol – $\Delta_f H^o$(298.15 K) = -601.576 kJ/mol – S^o(298.15 K) = 415.507 J/mol-K – p^o = 0.1 MPa



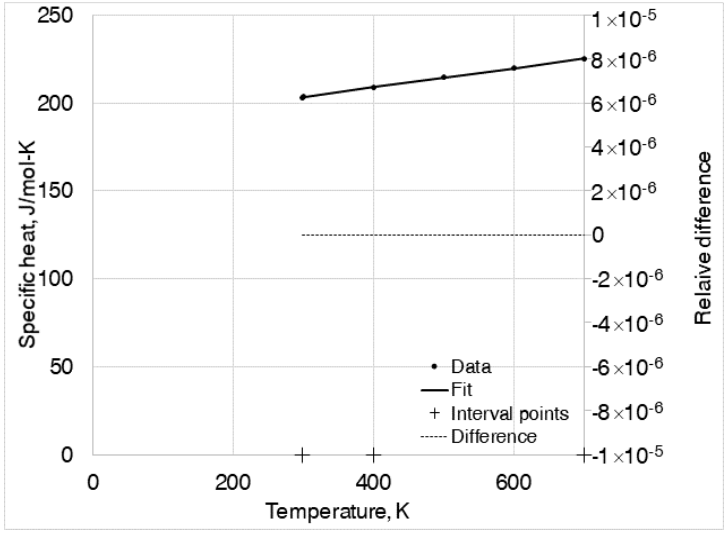


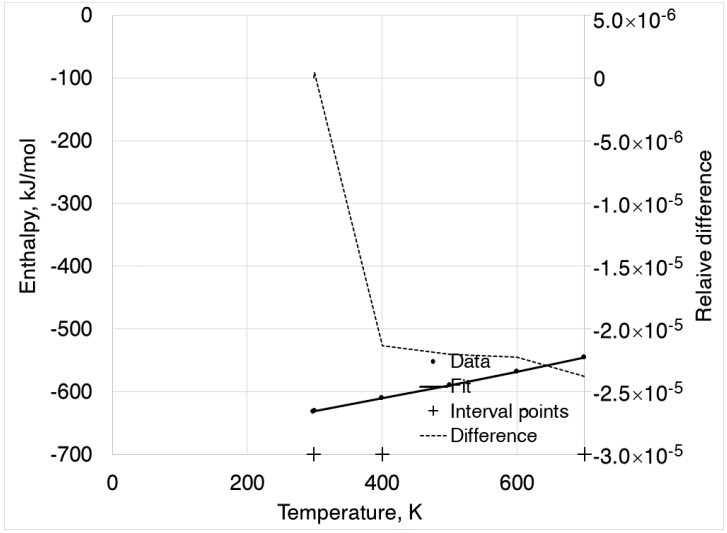


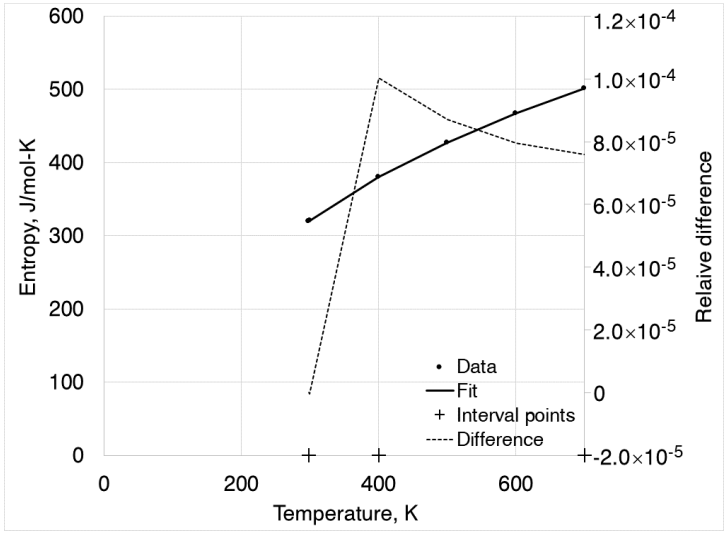


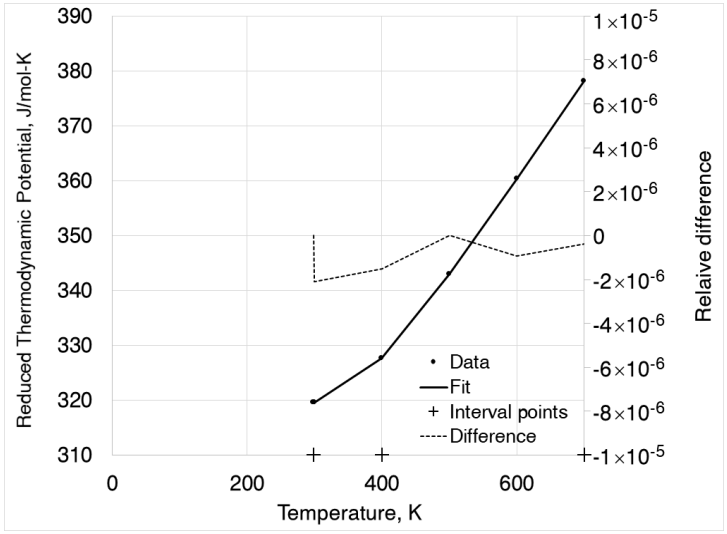
Nickel Carbonyl ($Ni(CO)_4$) – C4Ni1O4(l)

MW = 170.73 g/mol – $\Delta_f H^o$(298.15 K) = -631.784 kJ/mol – S^o(298.15 K) = 319.56 J/mol-K – p^o = 0.1 MPa



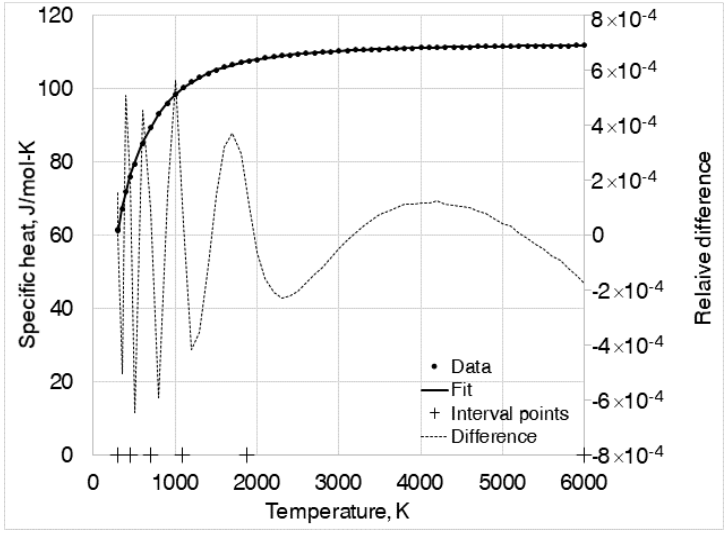


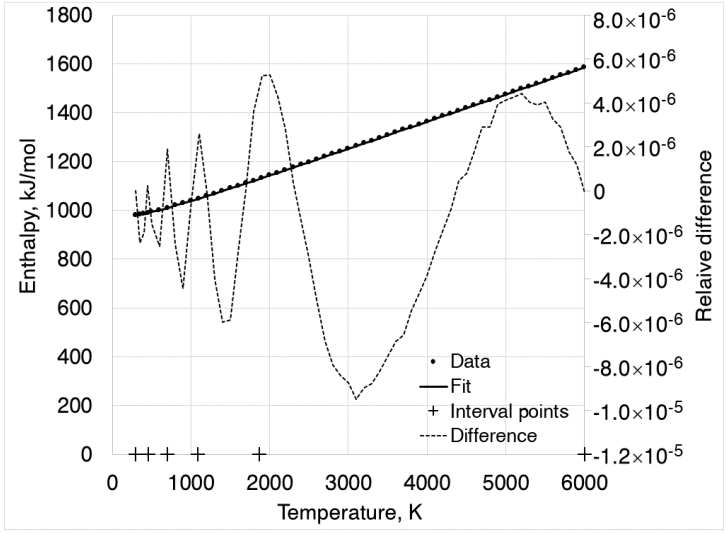


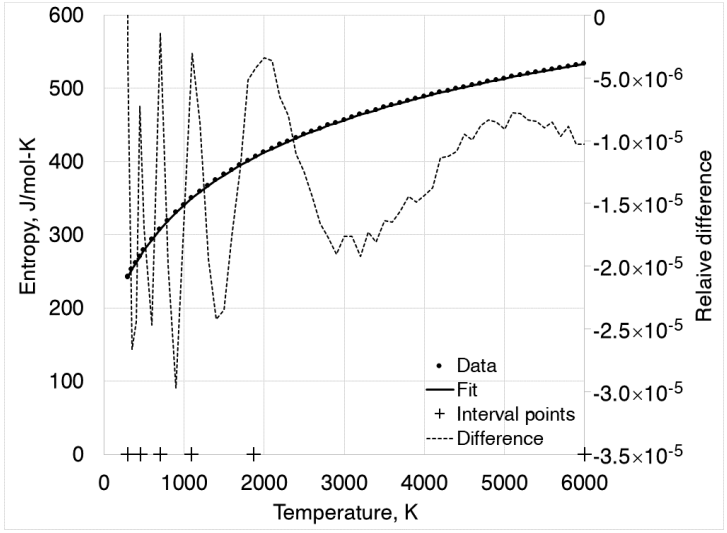


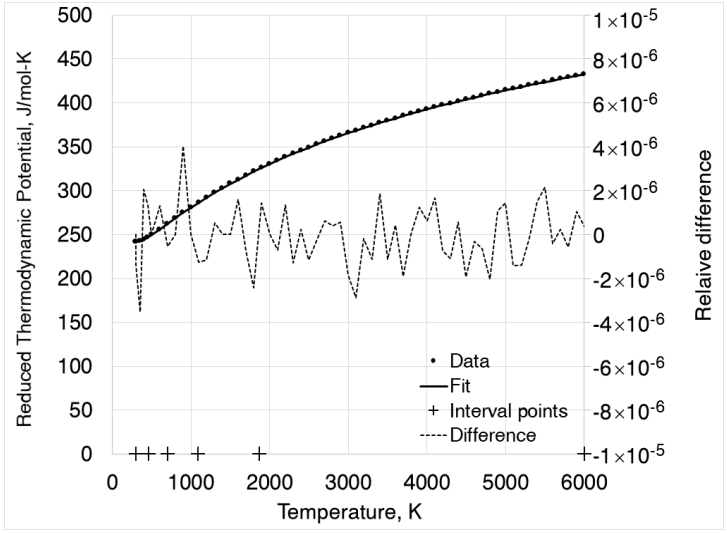
Carbon (C_5) – C5(g)

MW = 60.055 g/mol – $\Delta_f H^o$(298.15 K) = 979.056 kJ/mol – S^o(298.15 K) = 241.972 J/mol-K – p^o = 0.1 MPa



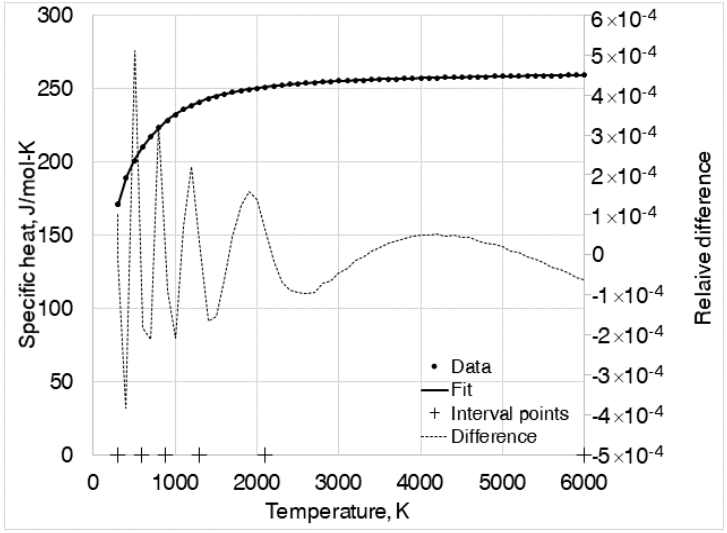


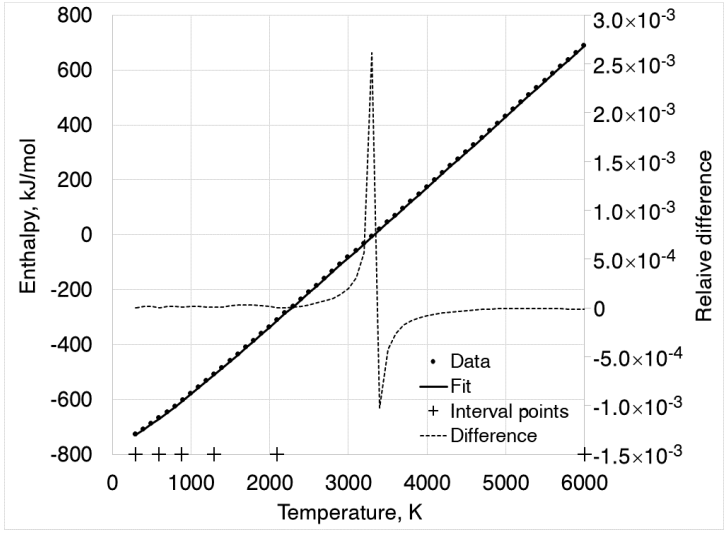


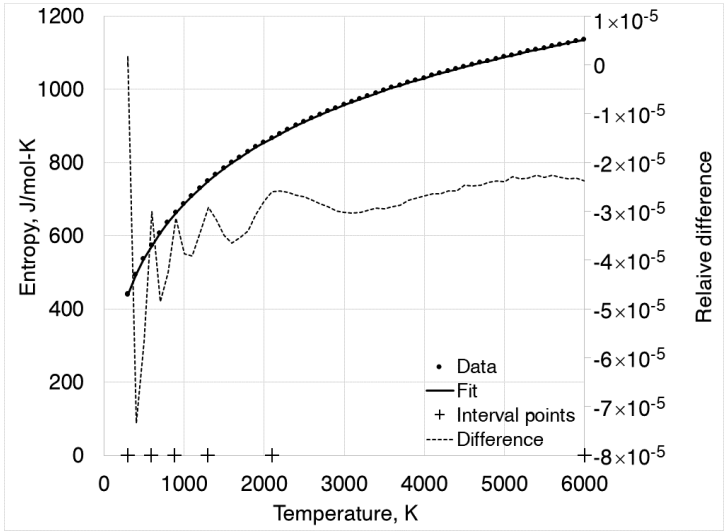


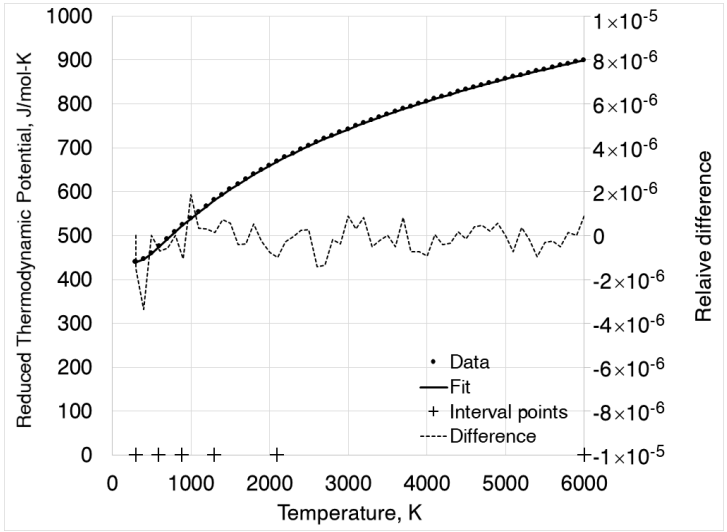
Iron Carbonyl ($Fe(CO)_5$) – C5Fe1O5(g)

MW = 195.899 g/mol – $\Delta_f H^o$(298.15 K) = -727.849 kJ/mol – S^o(298.15 K) = 439.286 J/mol-K – p^o = 0.1 MPa



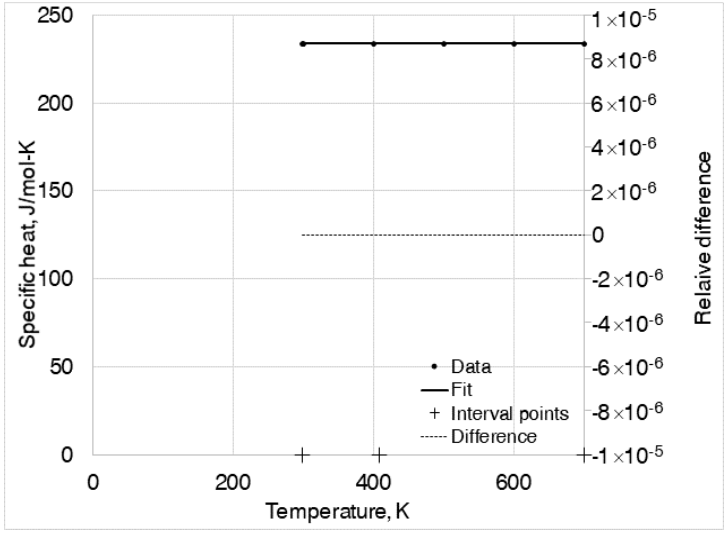


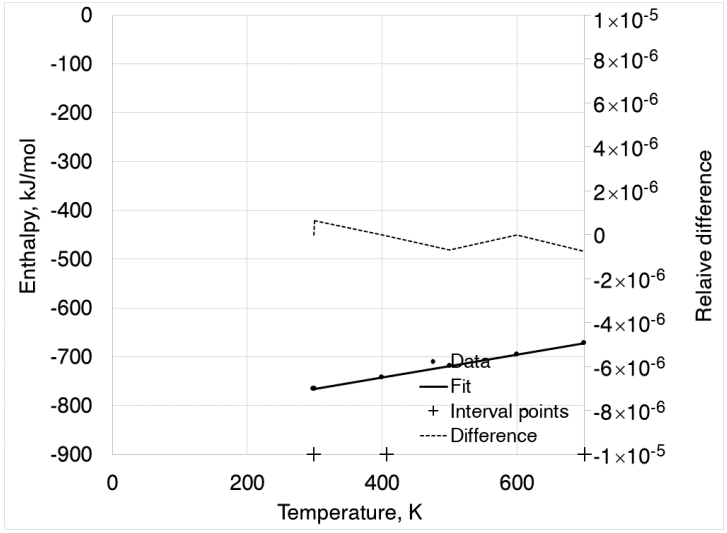


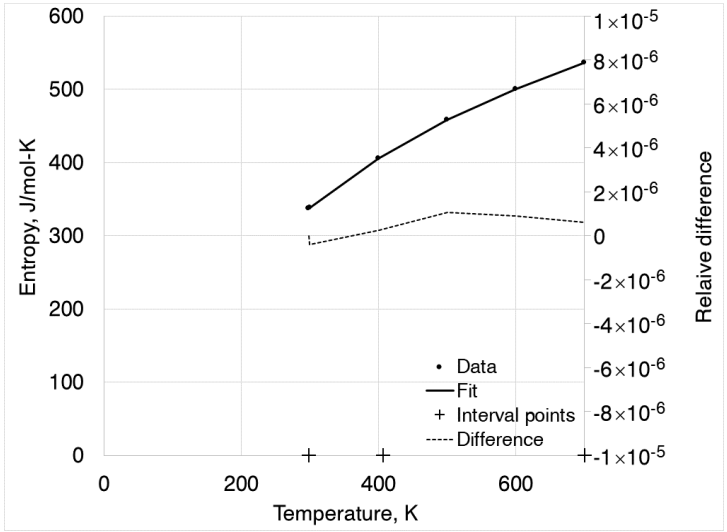


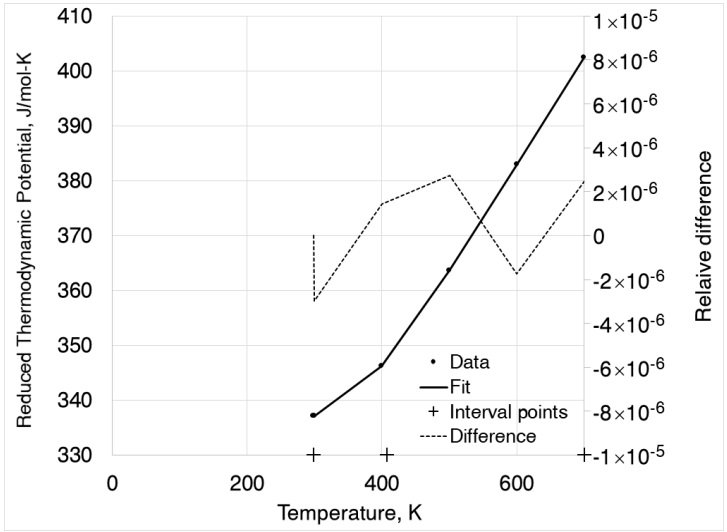
Iron Carbonyl ($Fe(CO)_5$) – C5Fe1O5(l)

MW = 195.899 g/mol – $\Delta_f H^o$(298.15 K) = -766.09 kJ/mol – S^o(298.15 K) = 337.078 J/mol-K – p^o = 0.1 MPa



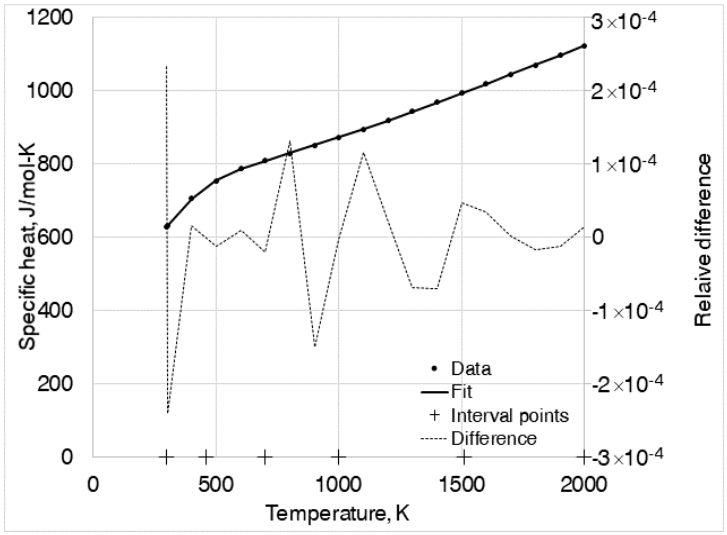


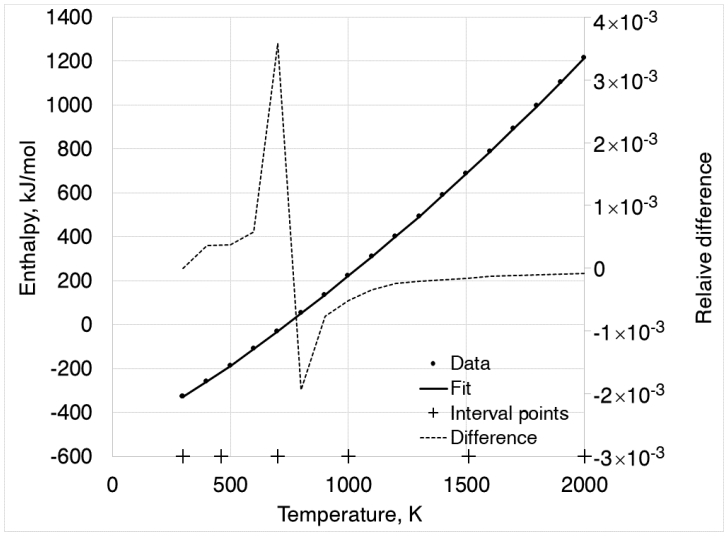


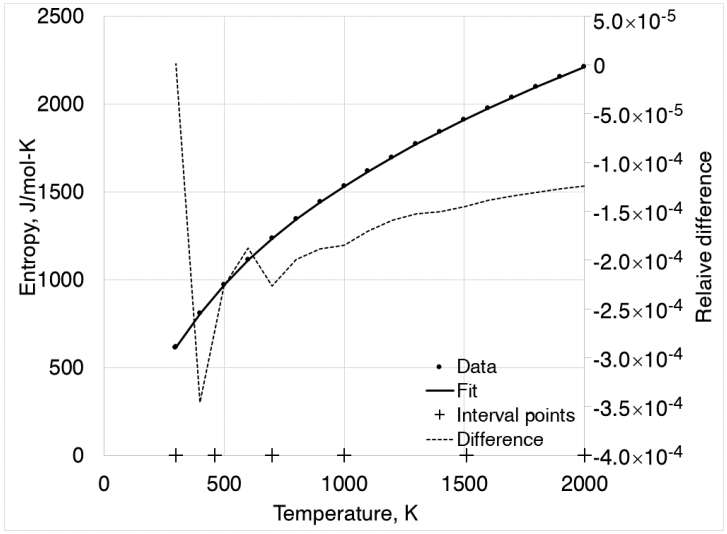


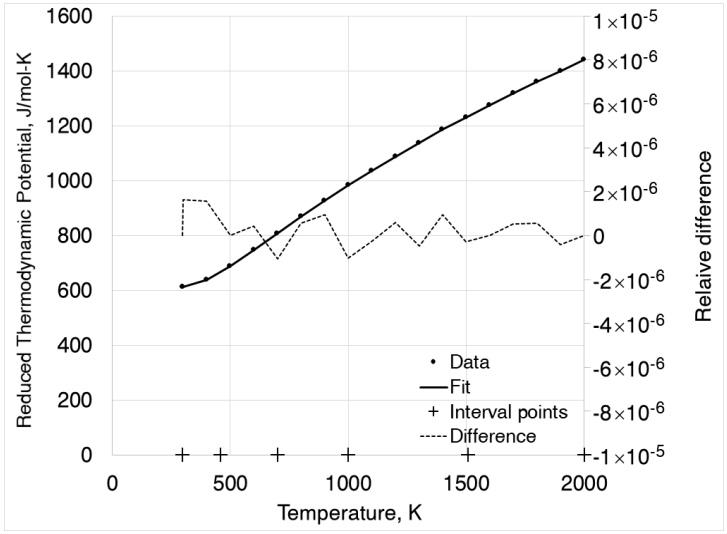
Chromium Carbide ($Cr_{23}C_6$) – C6Cr23(cr)

MW = 1267.974 g/mol – $\Delta_f H^o$(298.15 K) = -328.444 kJ/mol – S^o(298.15 K) = 612.044 J/mol-K – p^o = 0.1 MPa



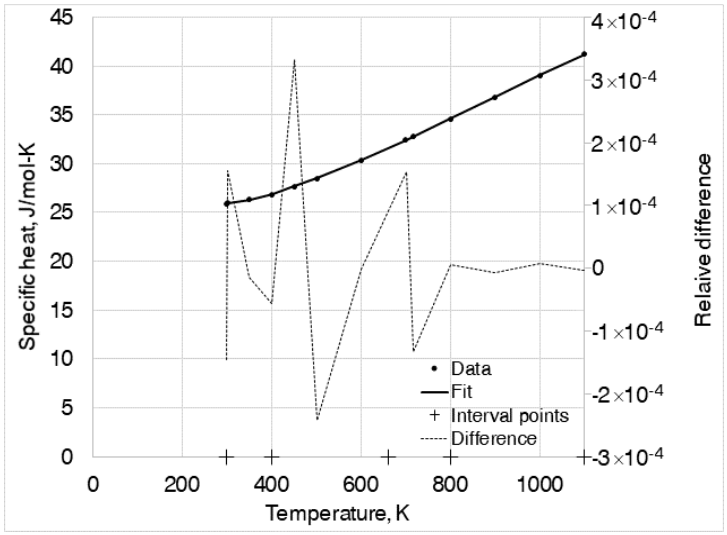


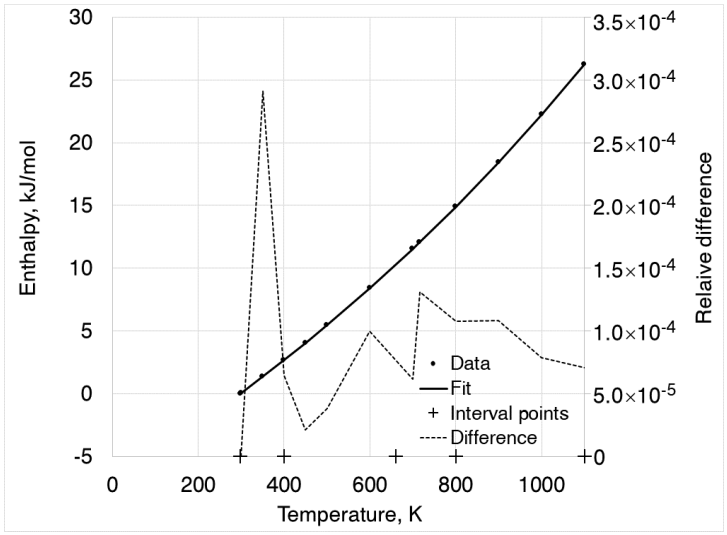


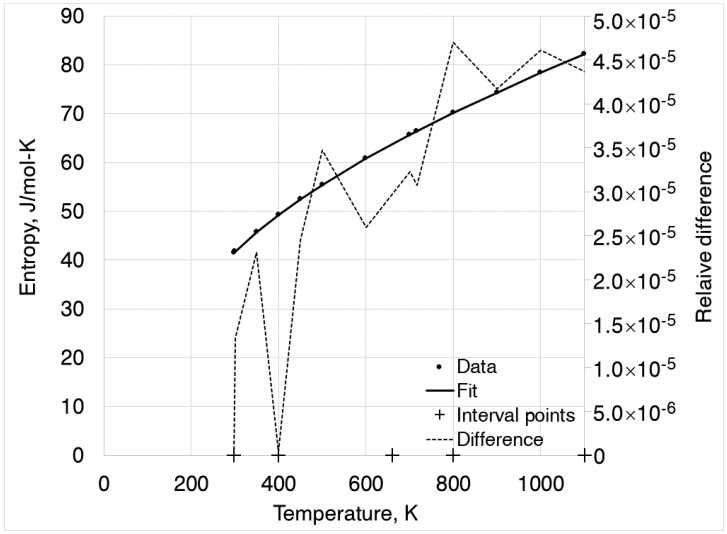


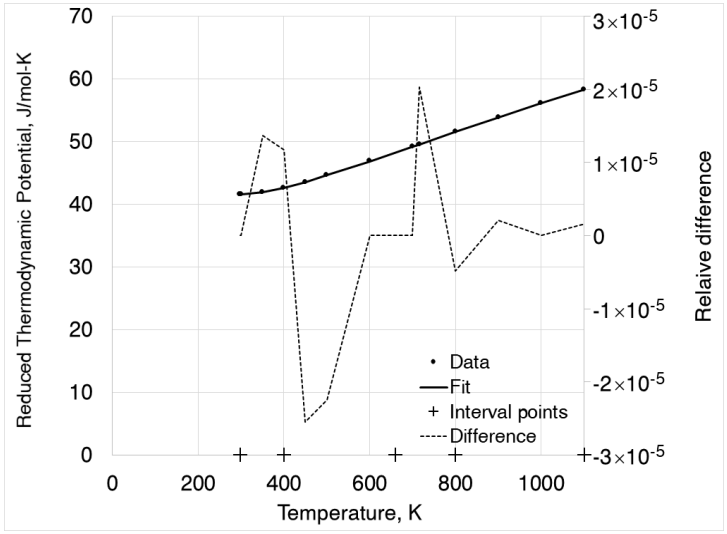
Calcium, Alpha (Ca) – Ca1(cr)_01

MW = 40.08 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 41.588 J/mol-K – p^o = 0.1 MPa



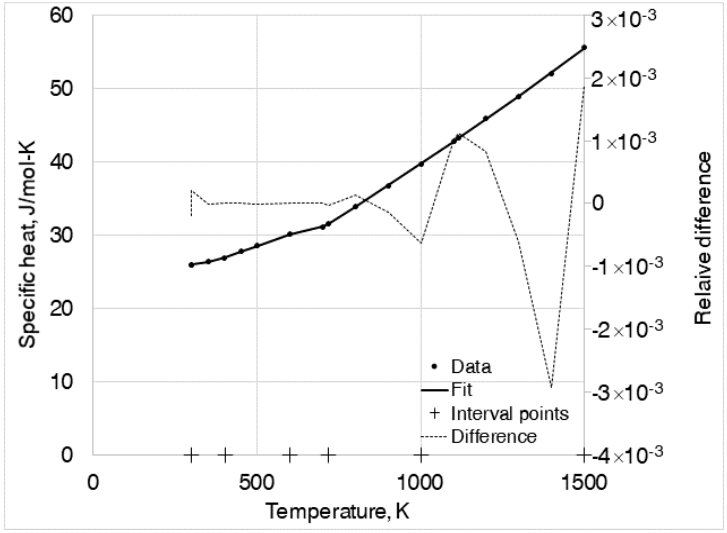


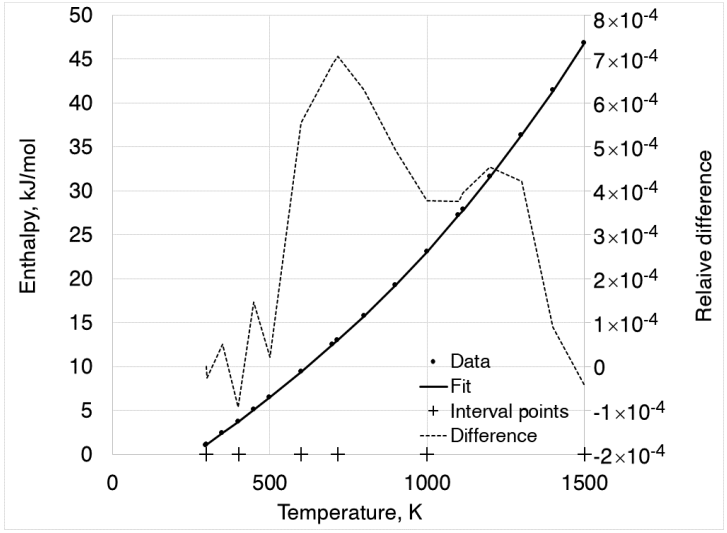


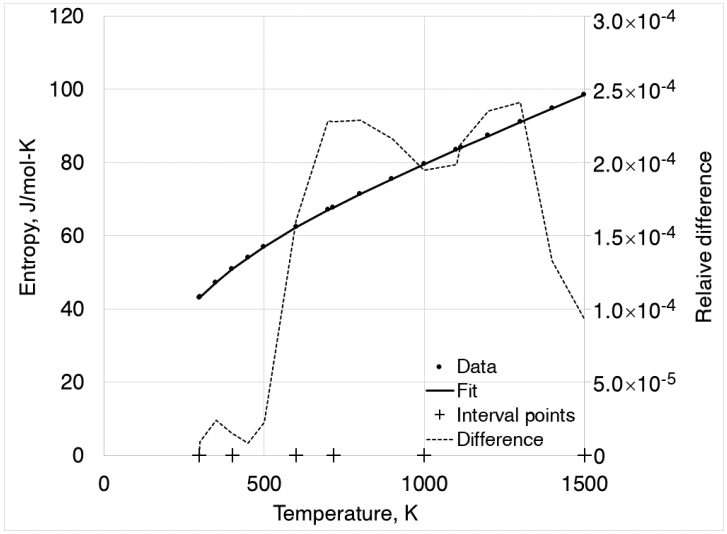


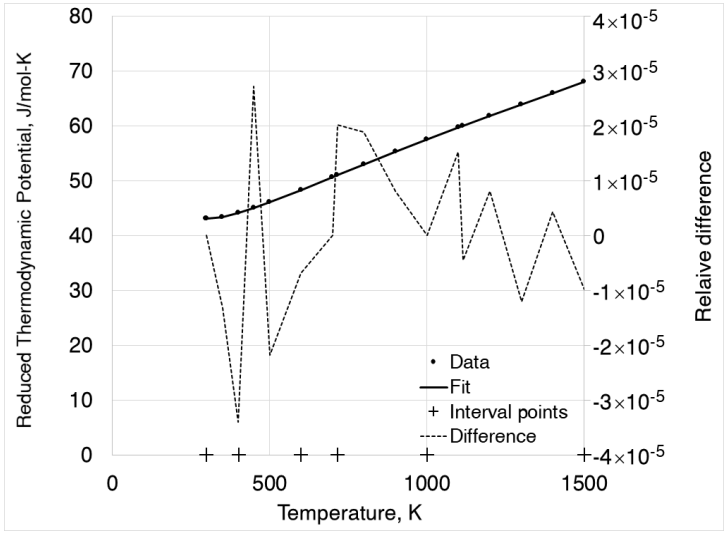
Calcium, Beta (Ca) – Ca1(cr)_02

MW = 40.08 g/mol – $\Delta_f H^o$(298.15 K) = 1.056 kJ/mol – S^o(298.15 K) = 43.07 J/mol-K – p^o = 0.1 MPa



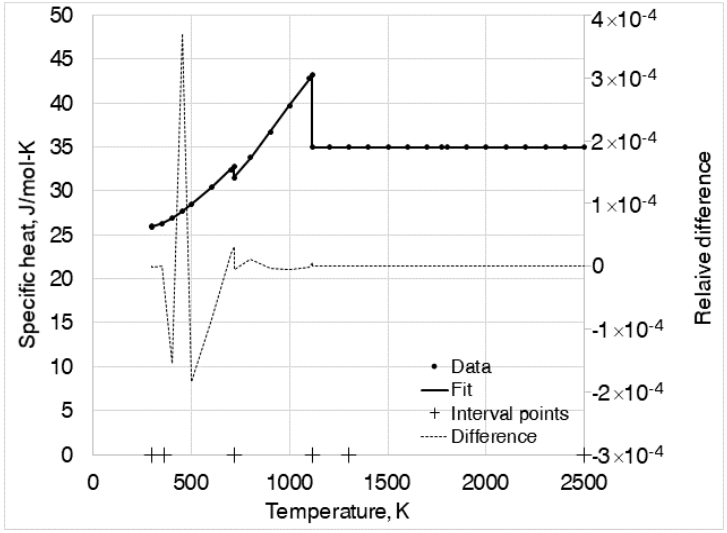


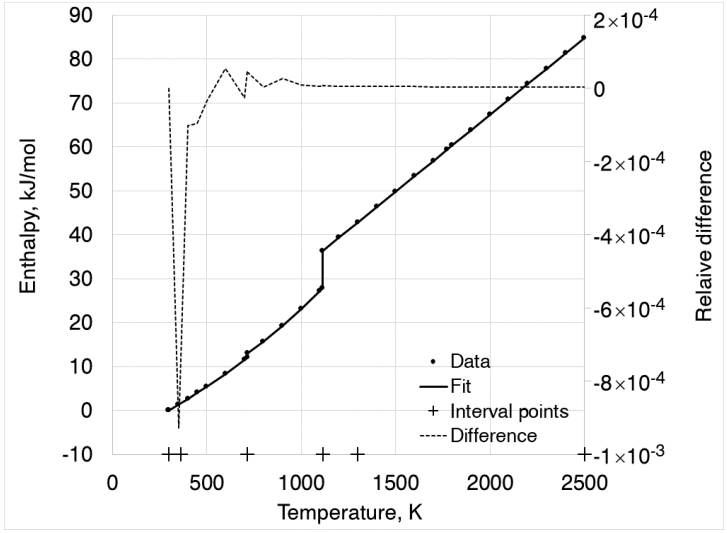


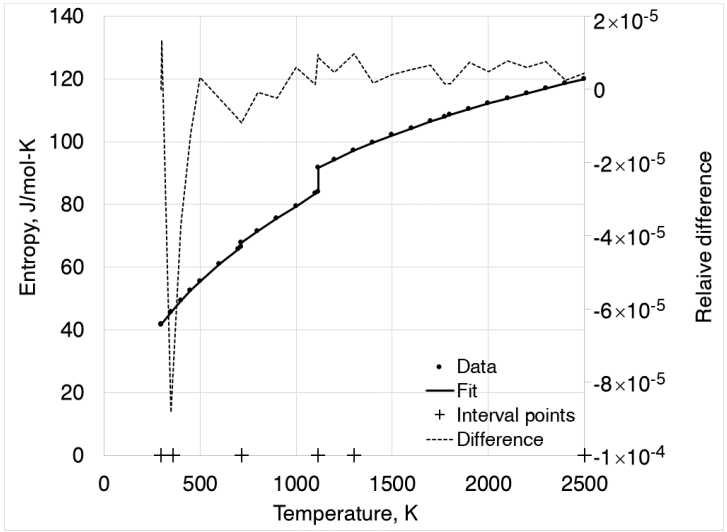


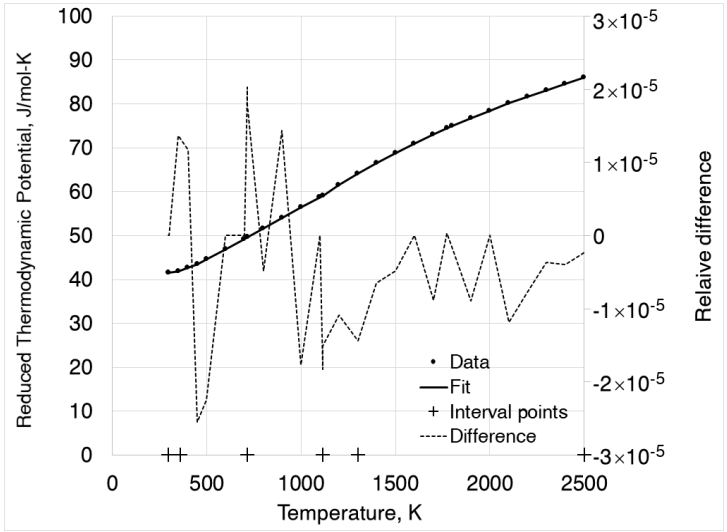
Calcium (Ca) – Ca1(cr,l)

MW = 40.08 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 41.588 J/mol-K – p^o = 0.1 MPa



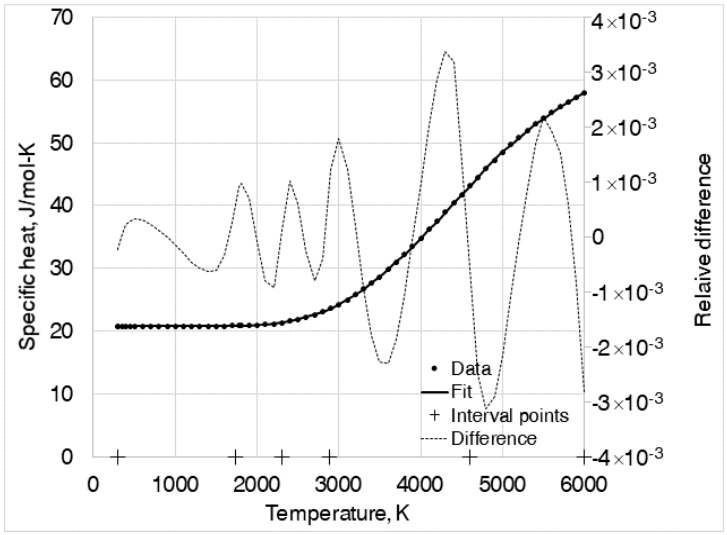


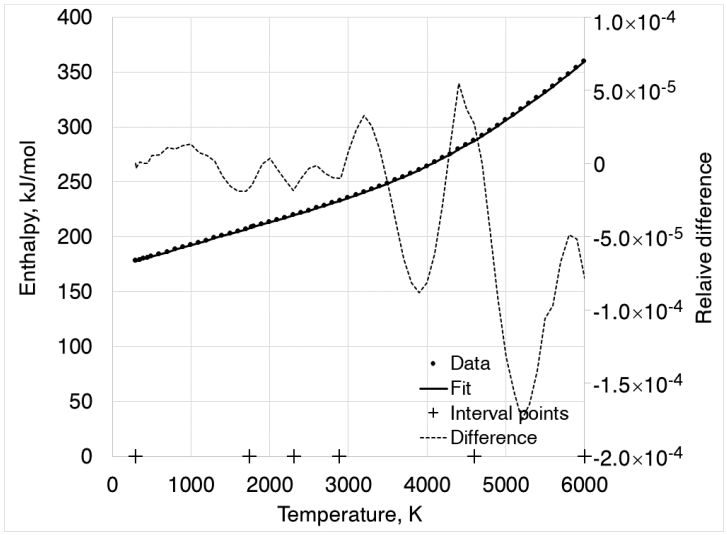


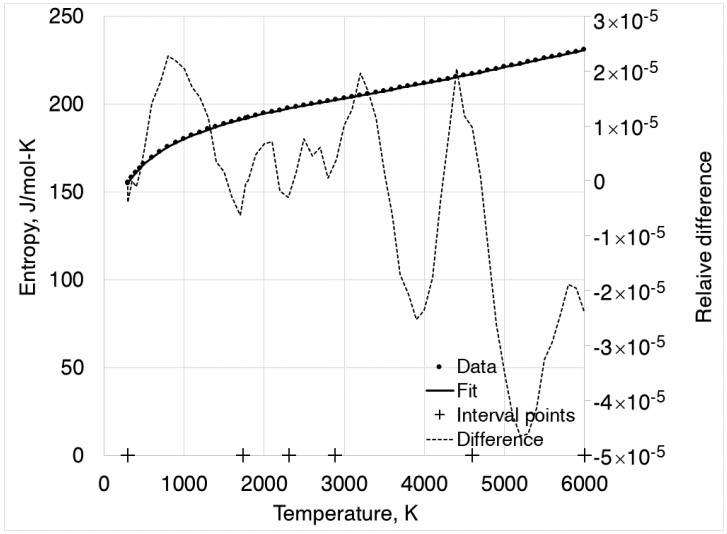


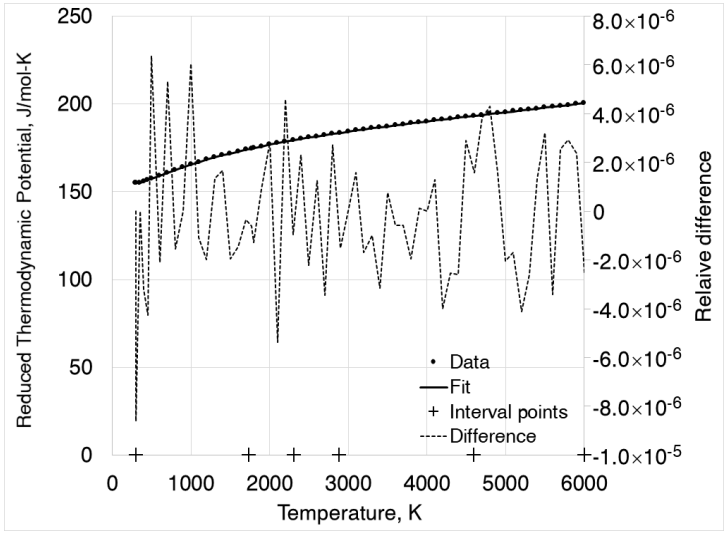
Calcium (Ca) – Ca1(g)

MW = 40.08 g/mol – $\Delta_f H^o$(298.15 K) = 177.8 kJ/mol – S^o(298.15 K) = 154.886 J/mol-K – p^o = 0.1 MPa



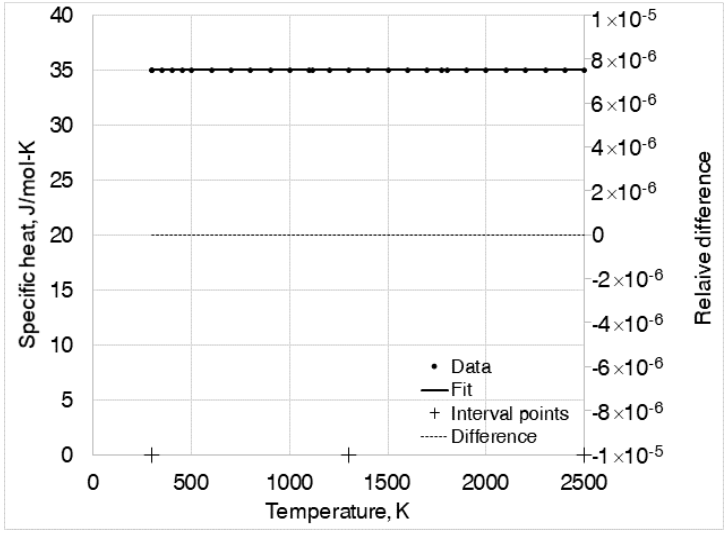


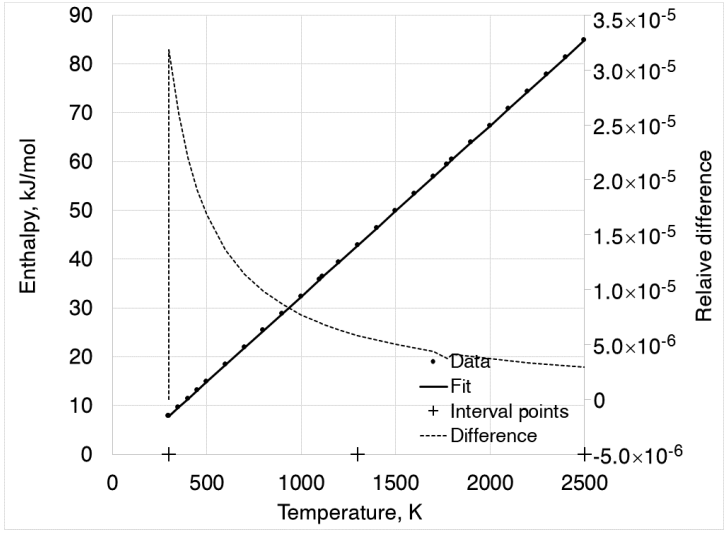


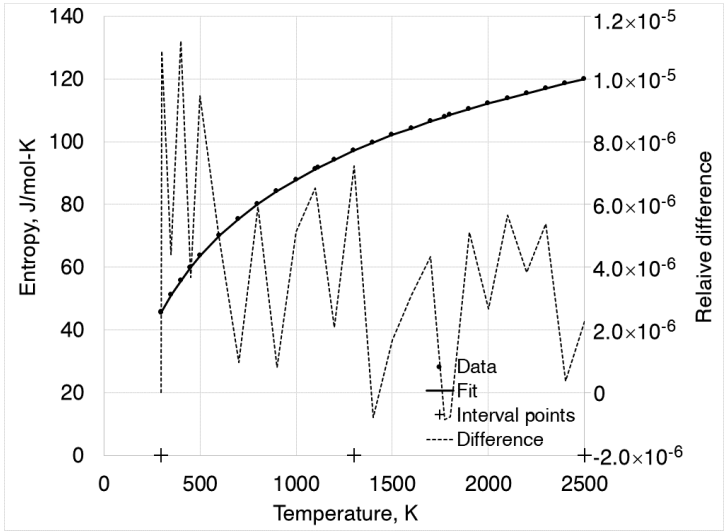


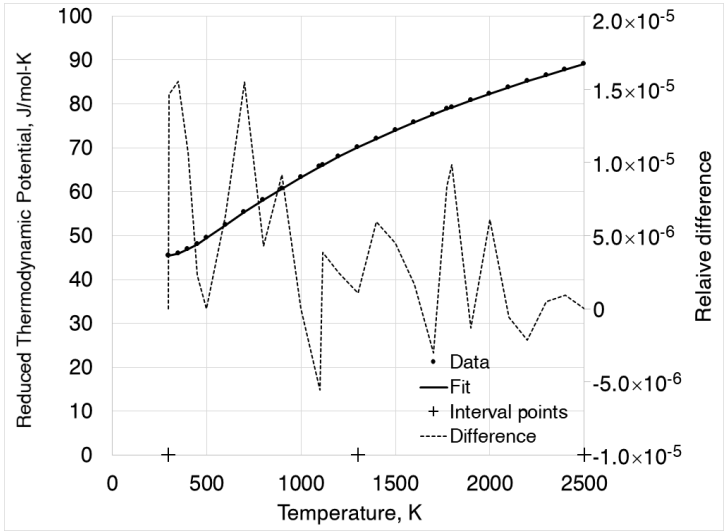
Calcium (Ca) – Ca1(l)

MW = 40.08 g/mol – $\Delta_f H^o$(298.15 K) = 7.788 kJ/mol – S^o(298.15 K) = 45.51 J/mol-K – p^o = 0.1 MPa



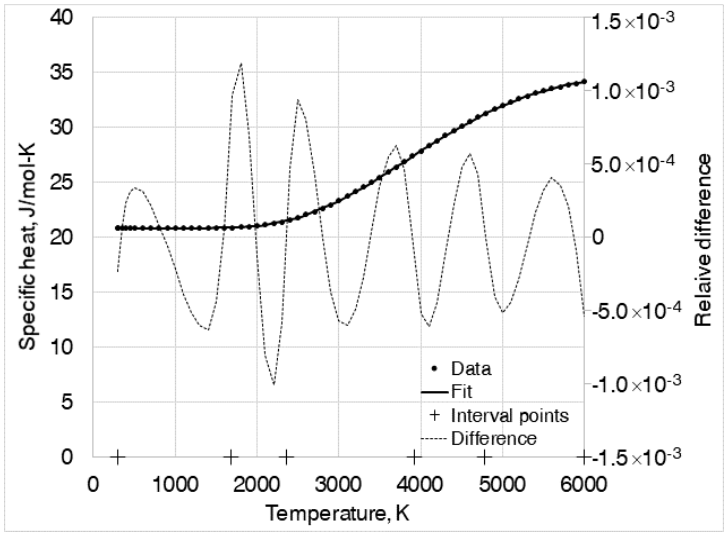


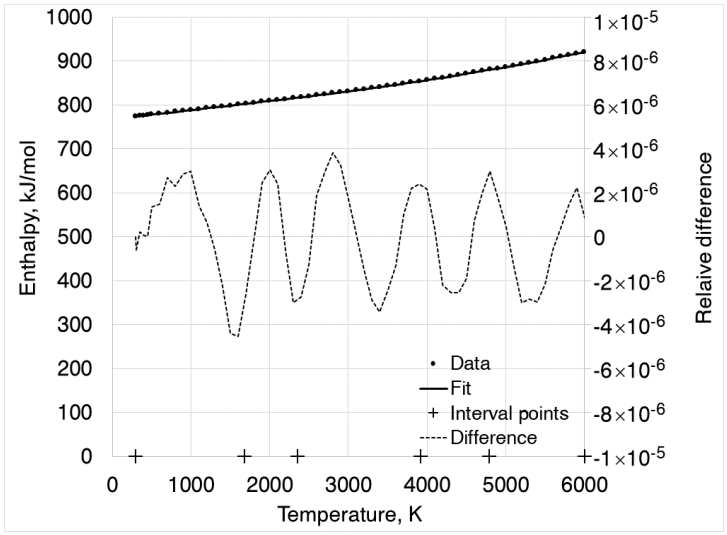


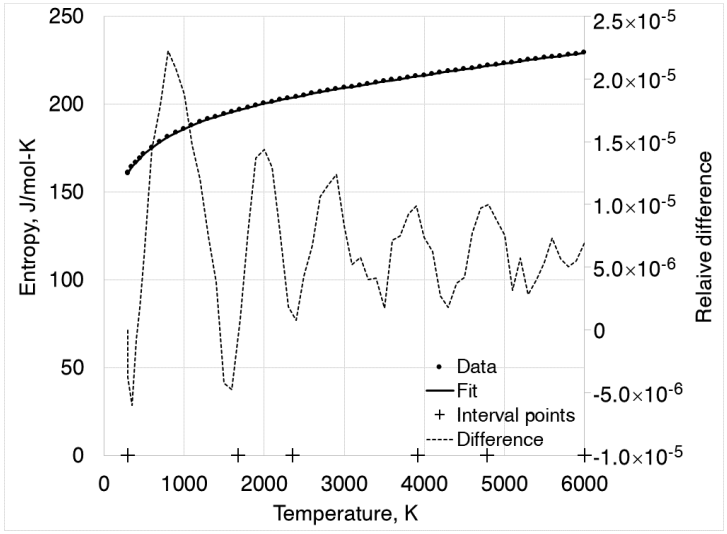


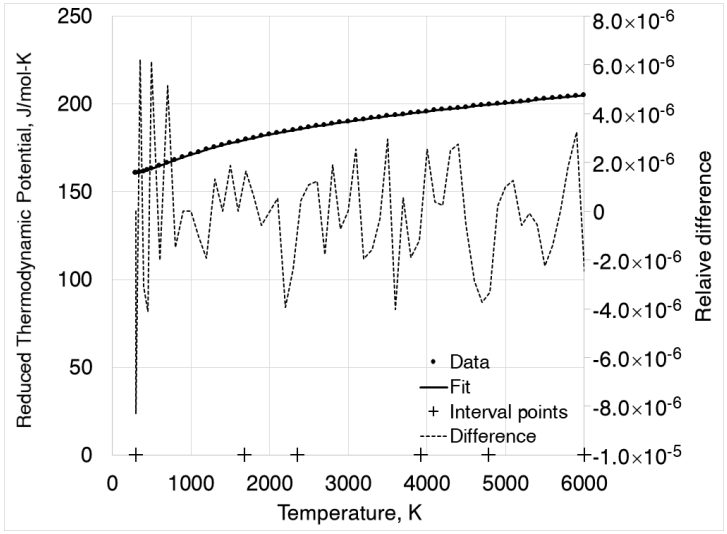
Calcium, Ion (Ca^+) – Ca1+(g)

MW = 40.08 g/mol – $\Delta_f H^o$(298.15 K) = 773.828 kJ/mol – S^o(298.15 K) = 160.649 J/mol-K – p^o = 0.1 MPa



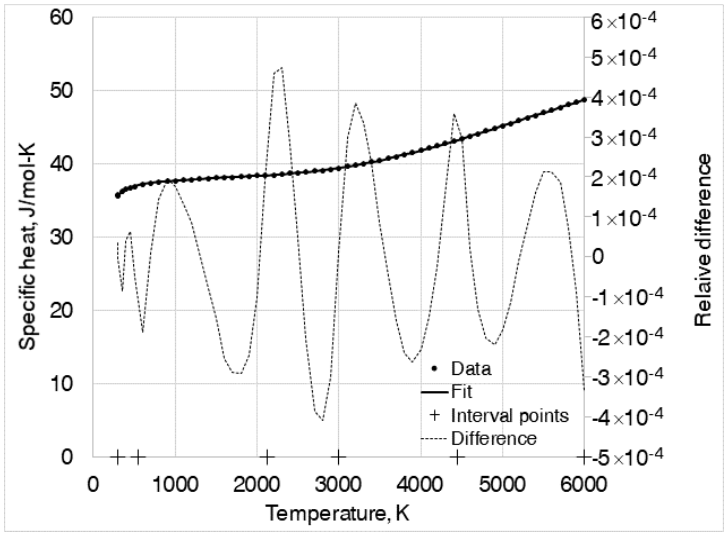


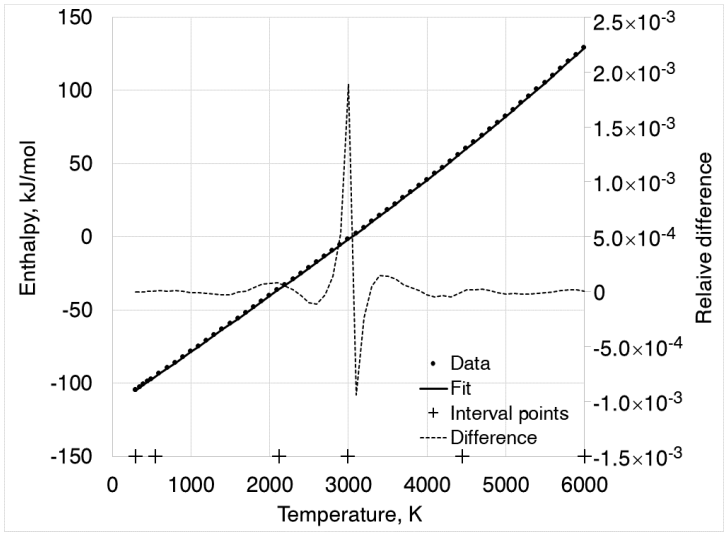


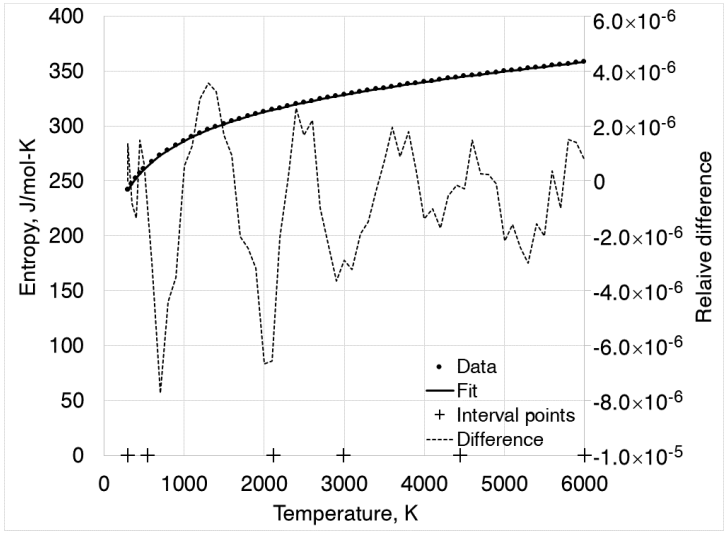


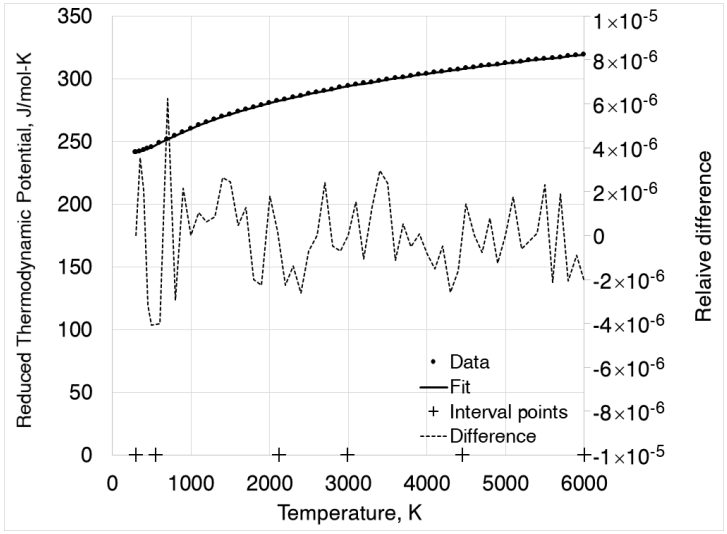
Calcium Chloride (CaCl) – Ca1Cl1(g)

MW = 75.53 g/mol – $\Delta_f H^o$(298.15 K) = -104.6 kJ/mol – S^o(298.15 K) = 241.559 J/mol-K – p^o = 0.1 MPa



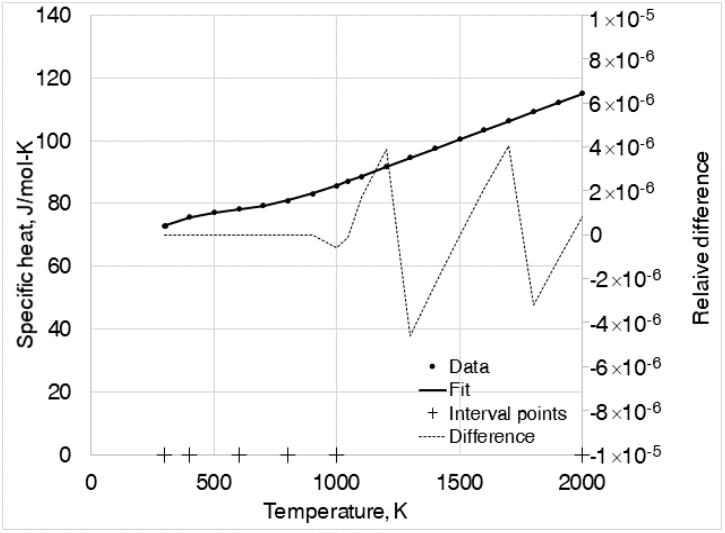


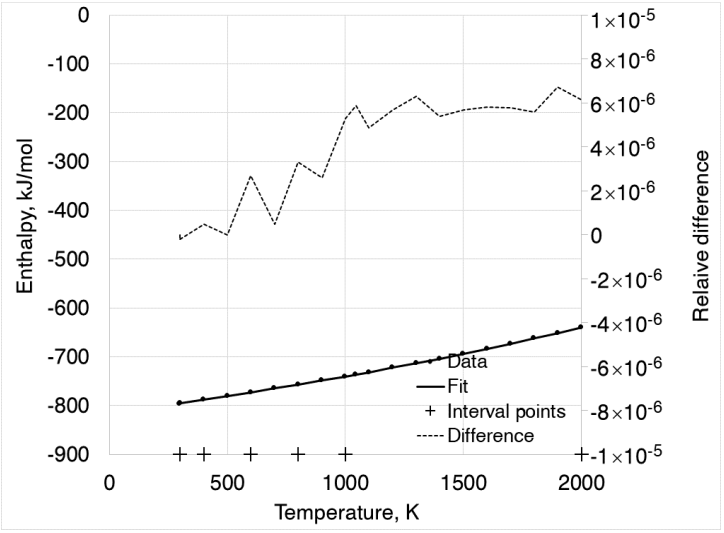


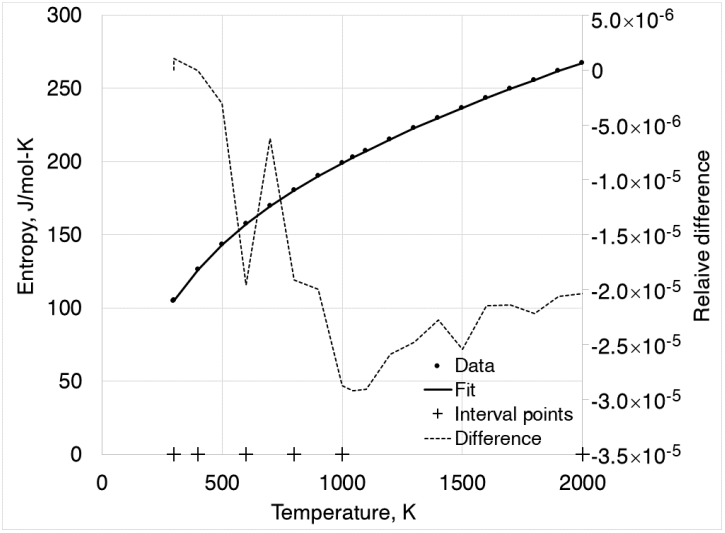


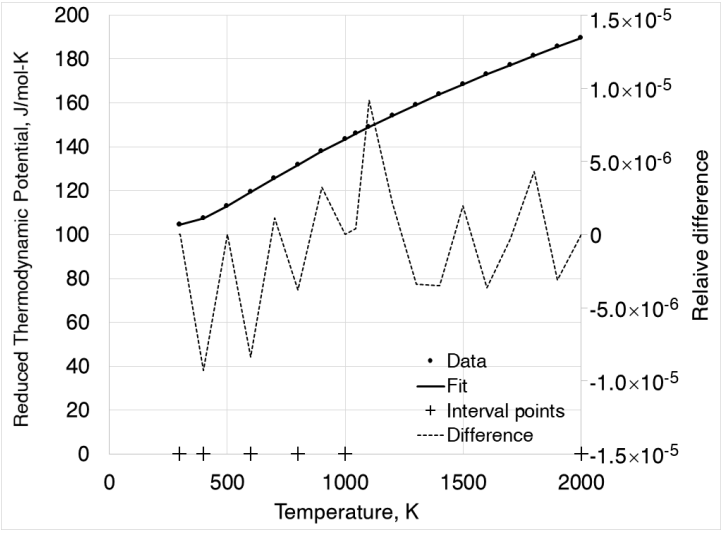
Calcium Chloride ($CaCl_2$) – Ca1Cl2(cr)

MW = 110.99 g/mol – $\Delta_f H^o$(298.15 K) = -795.797 kJ/mol – S^o(298.15 K) = 104.602 J/mol-K – p^o = 0.1 MPa



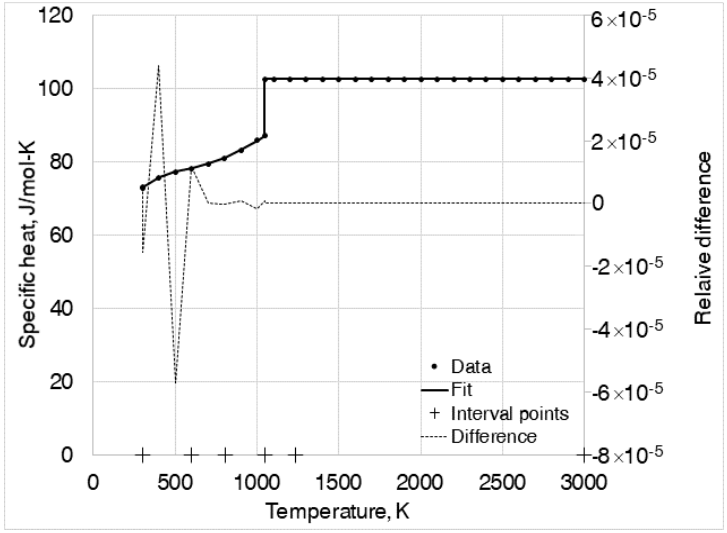


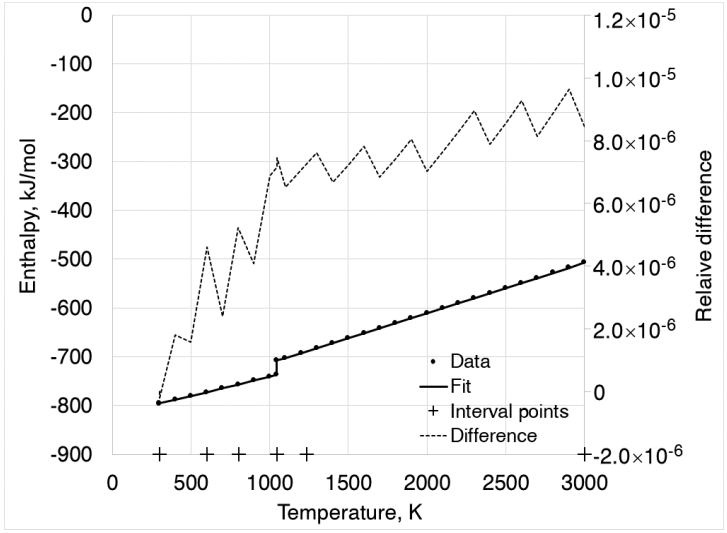


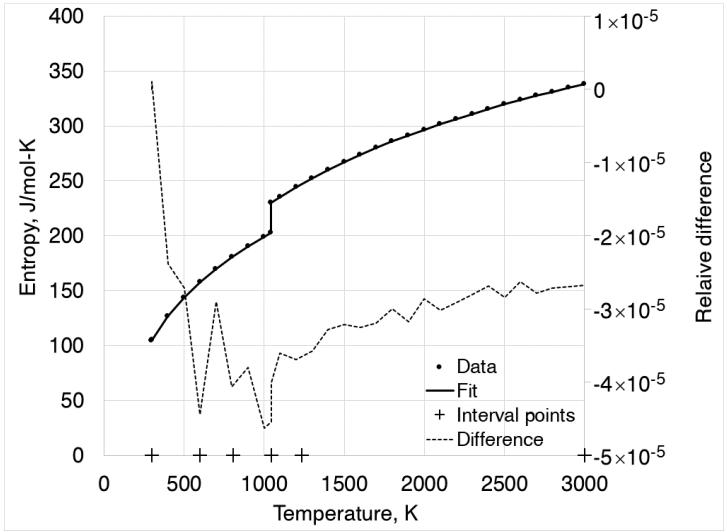


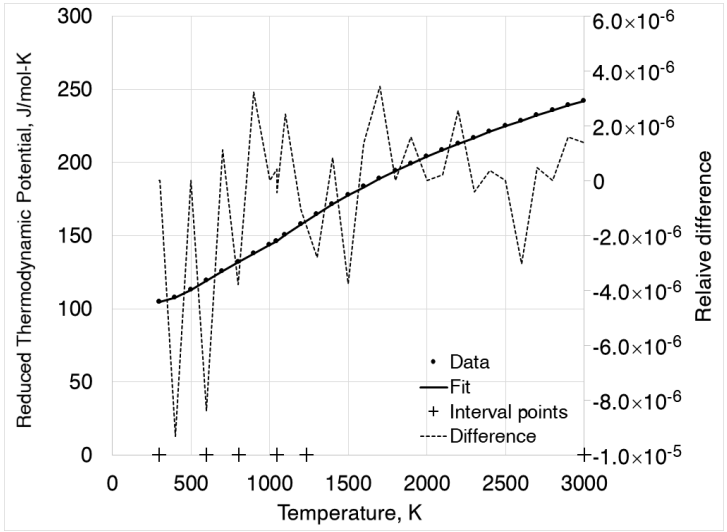
Calcium Chloride ($CaCl_2$) – Ca1Cl2(cr,l)

MW = 110.99 g/mol – $\Delta_f H^o$(298.15 K) = -795.797 kJ/mol – S^o(298.15 K) = 104.602 J/mol-K – p^o = 0.1 MPa



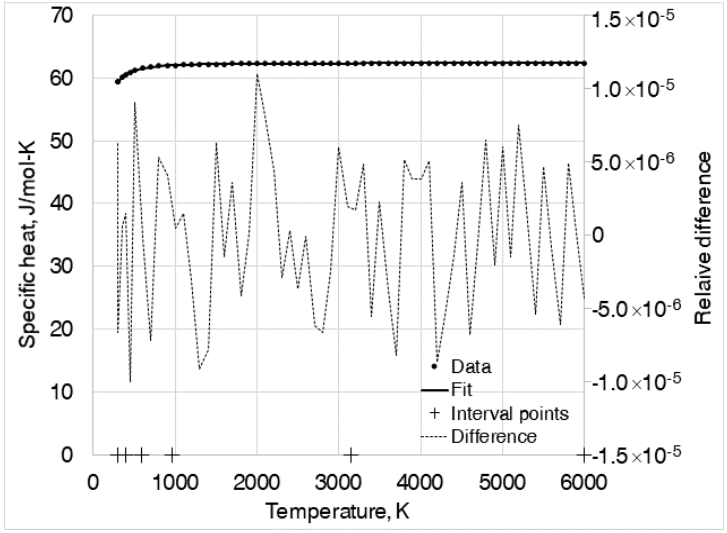


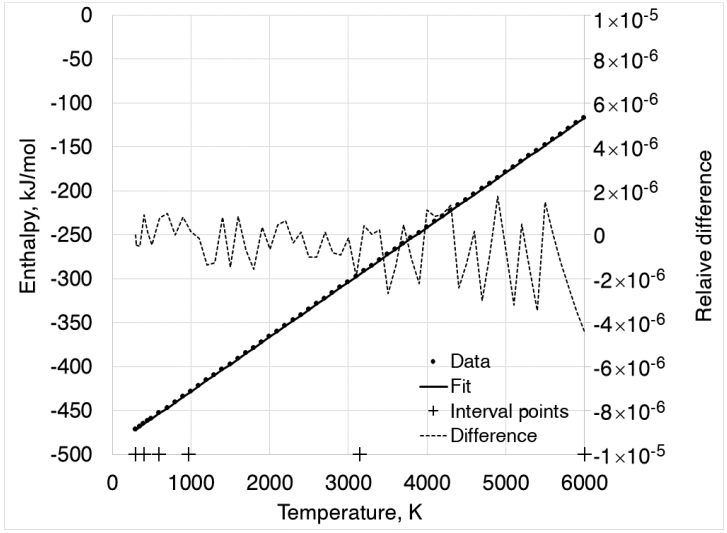


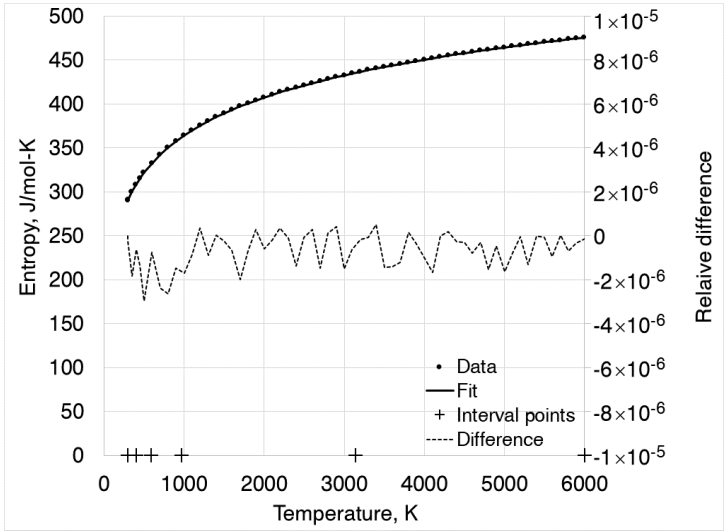


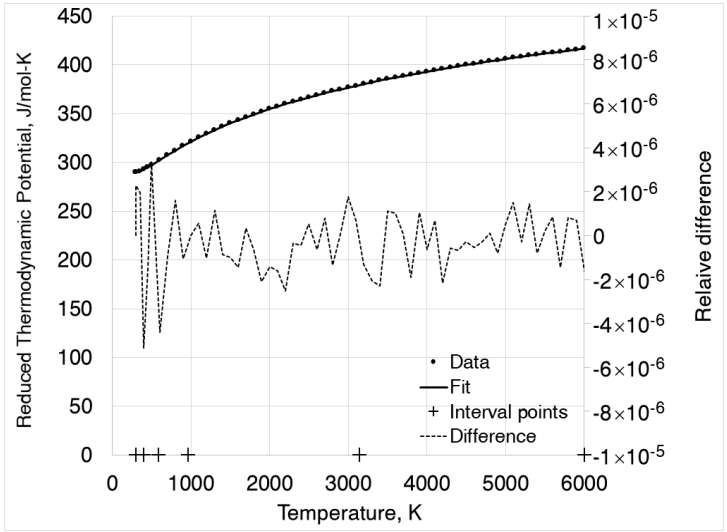
Calcium Chloride ($CaCl_2$) – Ca1Cl2(g)

MW = 110.99 g/mol – $\Delta_f H^o$(298.15 K) = -471.537 kJ/mol – S^o(298.15 K) = 290.265 J/mol-K – p^o = 0.1 MPa



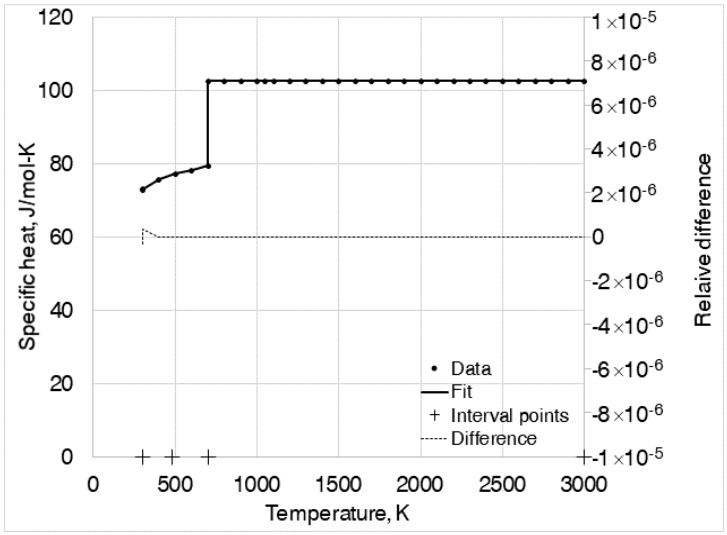


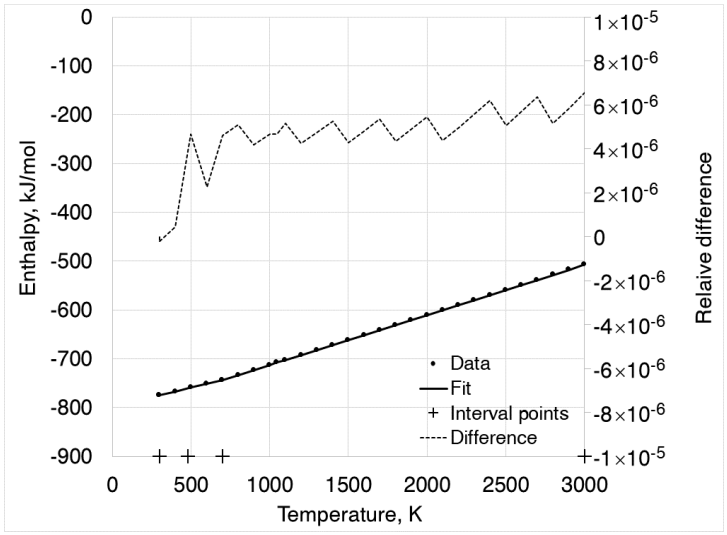


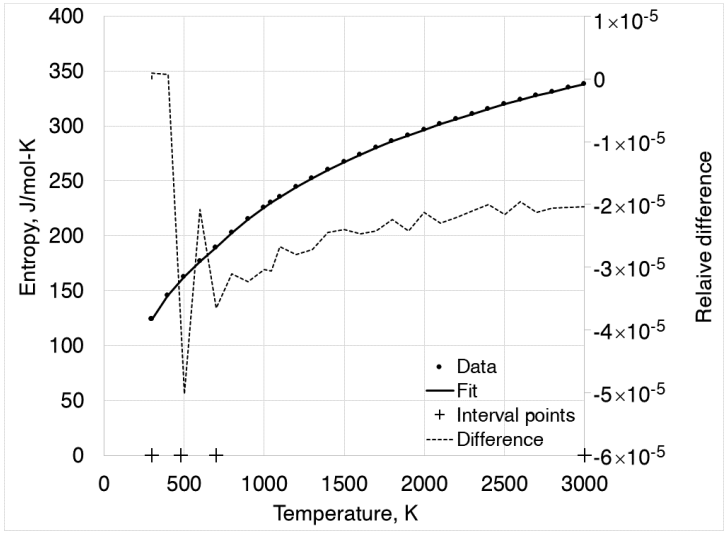


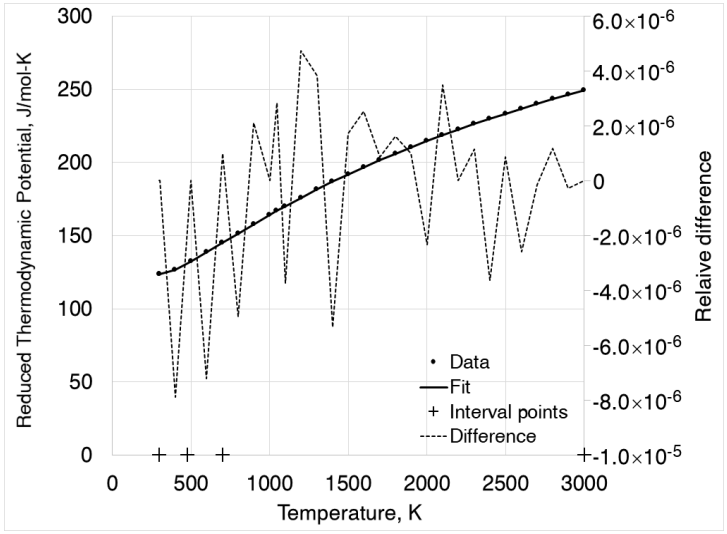
Calcium Chloride ($CaCl_2$) – Ca1Cl2(l)

MW = 110.99 g/mol – $\Delta_f H^o$(298.15 K) = -774.094 kJ/mol – S^o(298.15 K) = 123.868 J/mol-K – p^o = 0.1 MPa



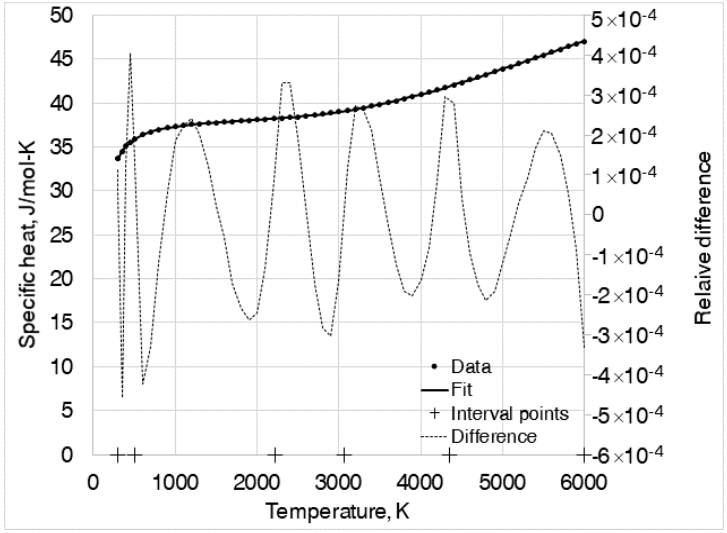


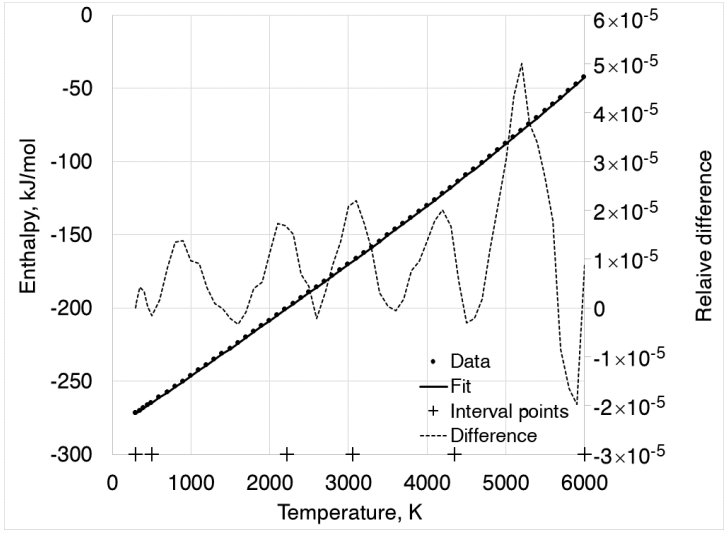


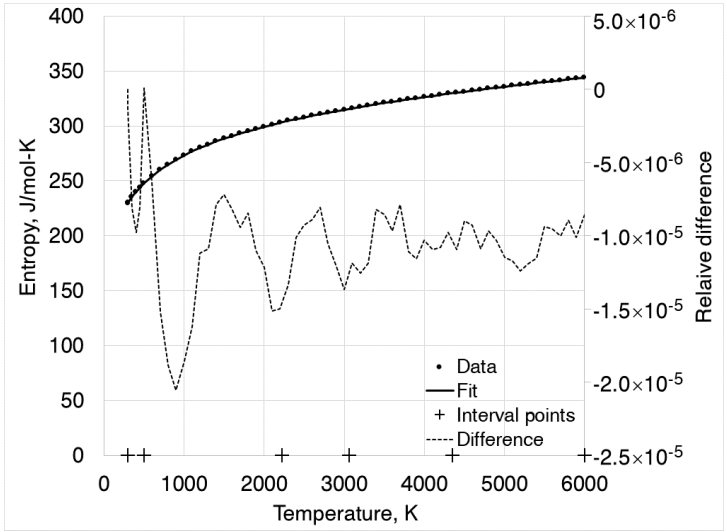


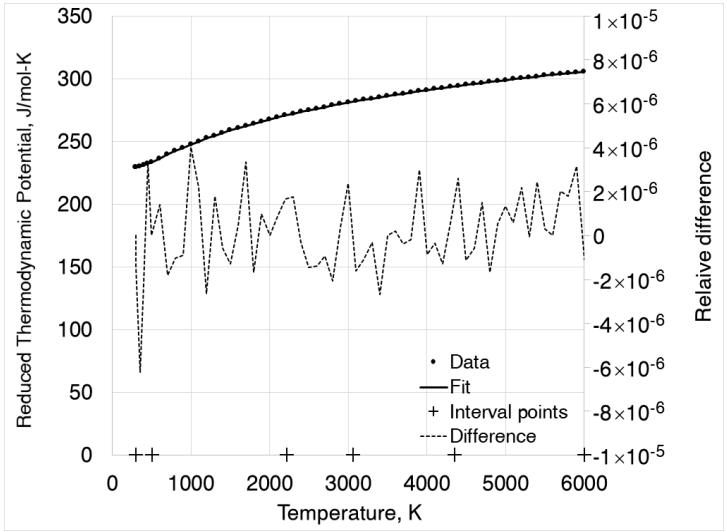
Calcium Fluoride (CaF) – Ca1F1(g)

MW = 59.08 g/mol – $\Delta_f H^o$(298.15 K) = -271.96 kJ/mol – S^o(298.15 K) = 229.647 J/mol-K – p^o = 0.1 MPa



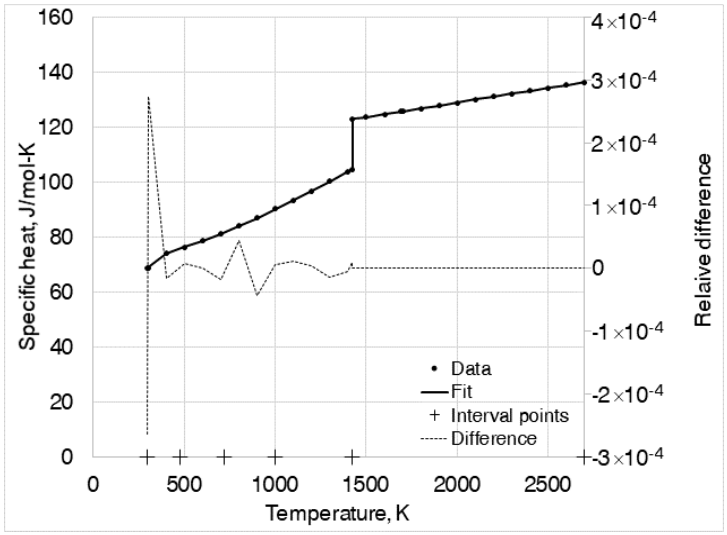


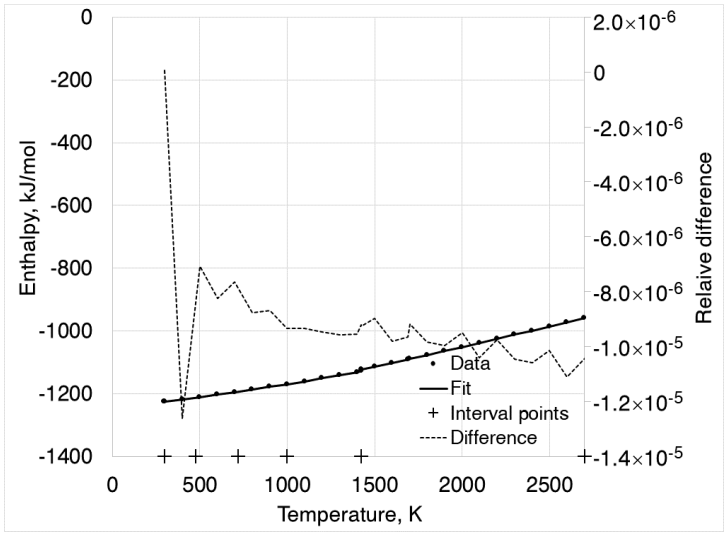


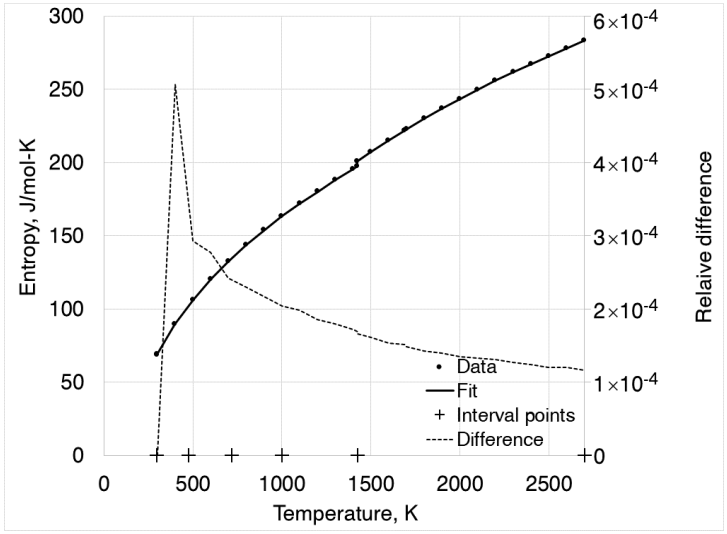


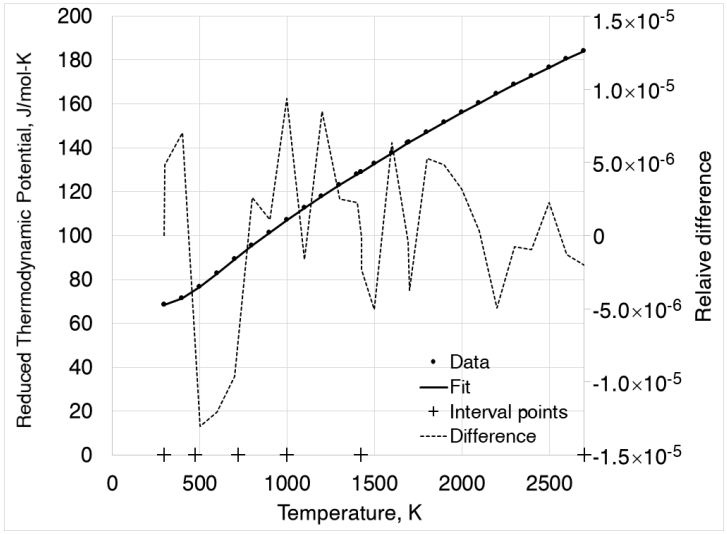
Calcium Fluoride (CaF_2) – Ca1F2(cr)

MW = 78.08 g/mol – $\Delta_f H^o$(298.15 K) = -1225.912 kJ/mol – S^o(298.15 K) = 68.572 J/mol-K – p^o = 0.1 MPa



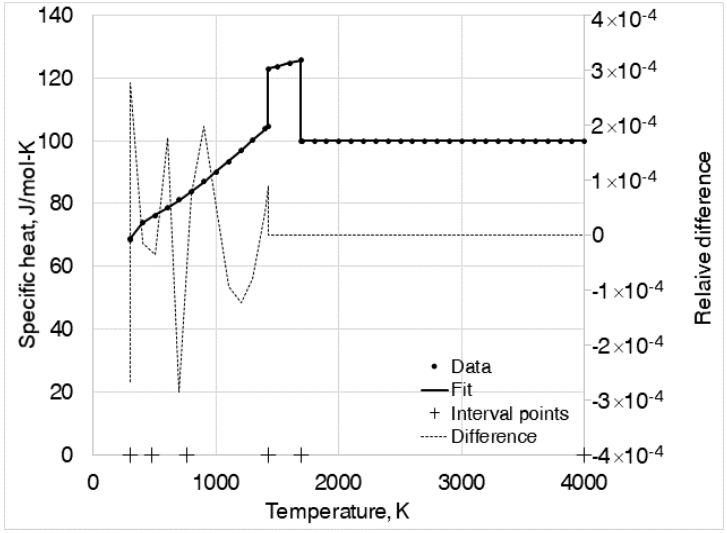


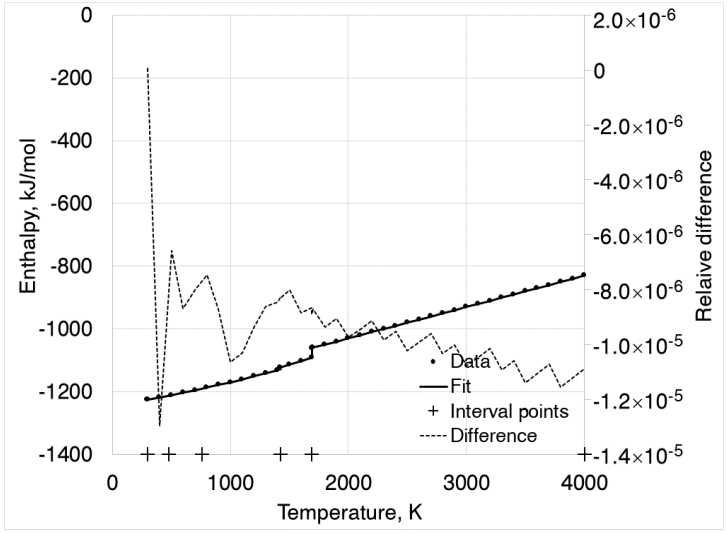


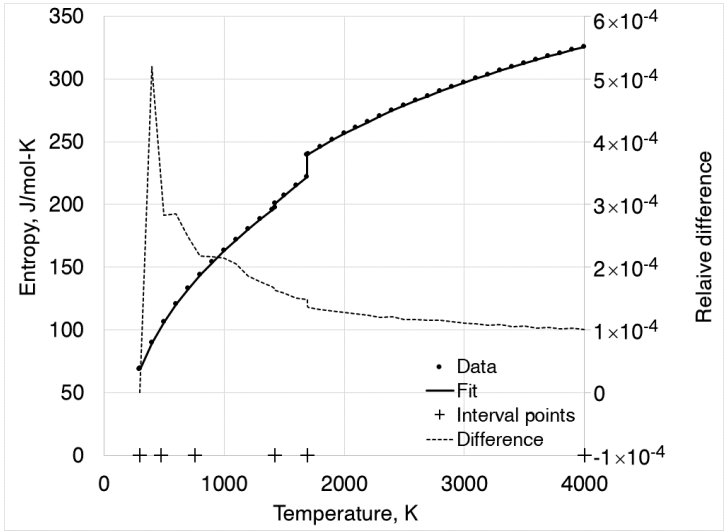


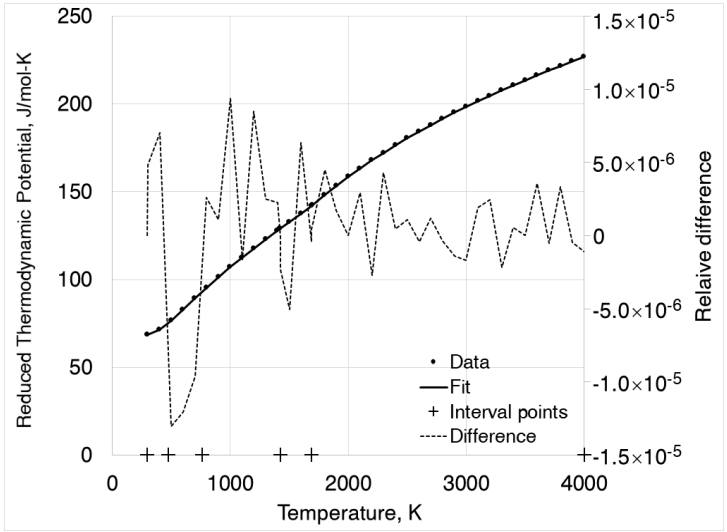
Calcium Fluoride (CaF_2) – Ca1F2(cr,l)

MW = 78.08 g/mol – $\Delta_f H^o$(298.15 K) = -1225.912 kJ/mol – S^o(298.15 K) = 68.572 J/mol-K – p^o = 0.1 MPa



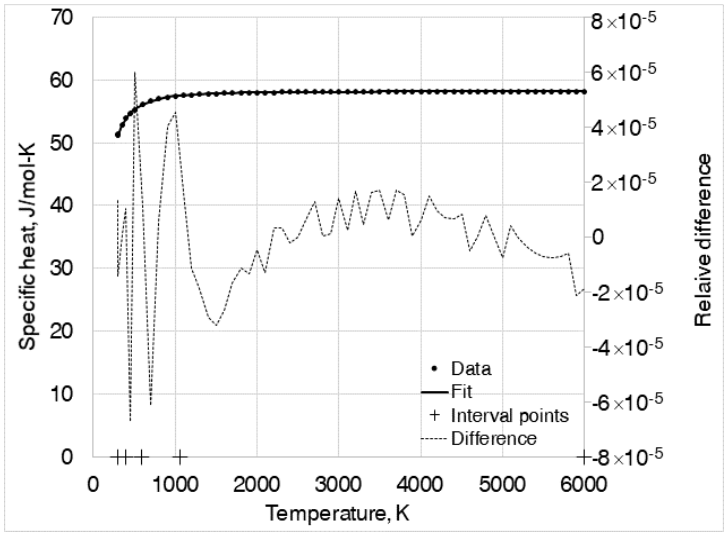


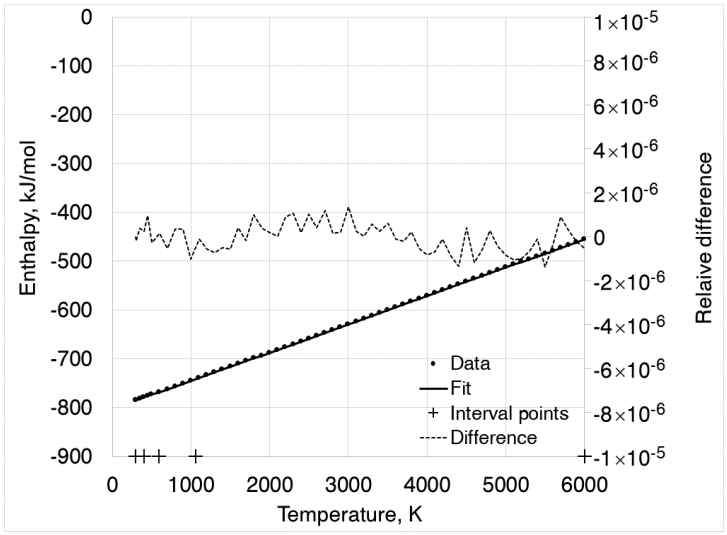


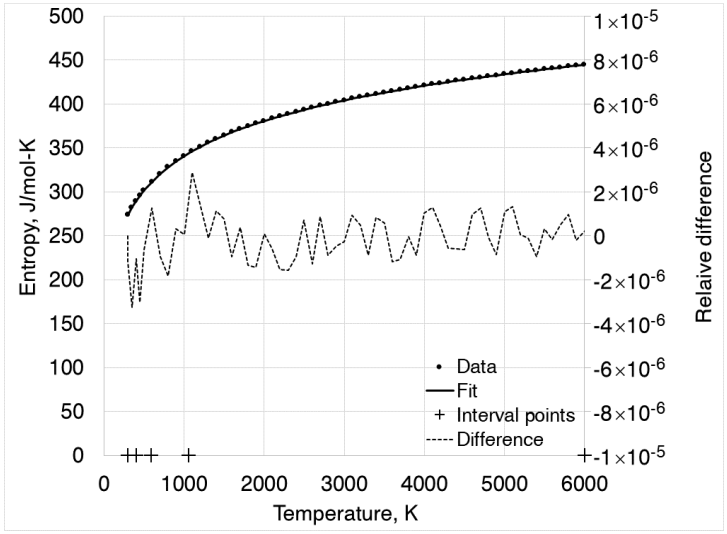


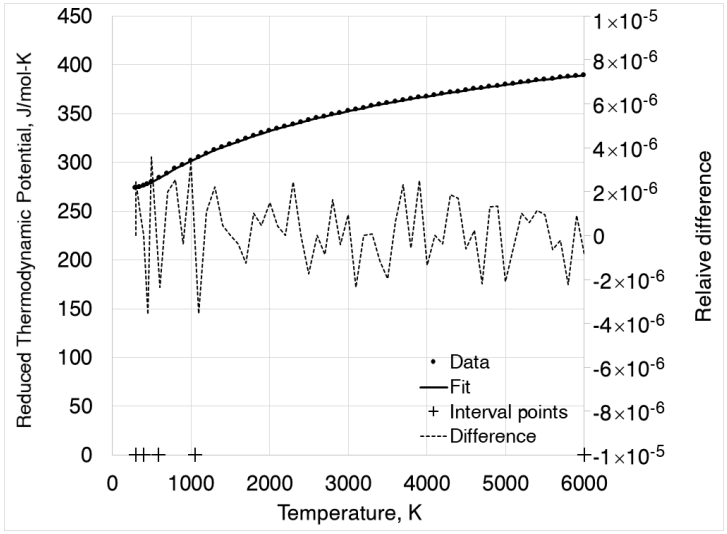
Calcium Fluoride (CaF_2) – Ca1F2(g)

MW = 78.08 g/mol – $\Delta_f H^o$(298.15 K) = -784.5 kJ/mol – S^o(298.15 K) = 273.798 J/mol-K – p^o = 0.1 MPa



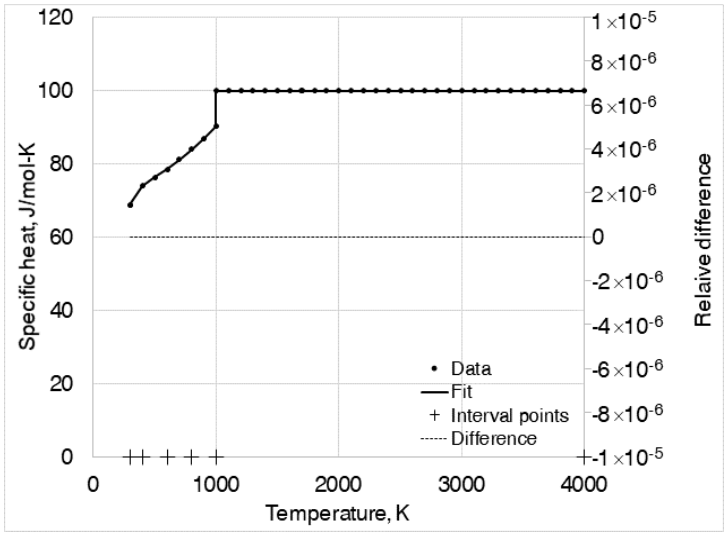


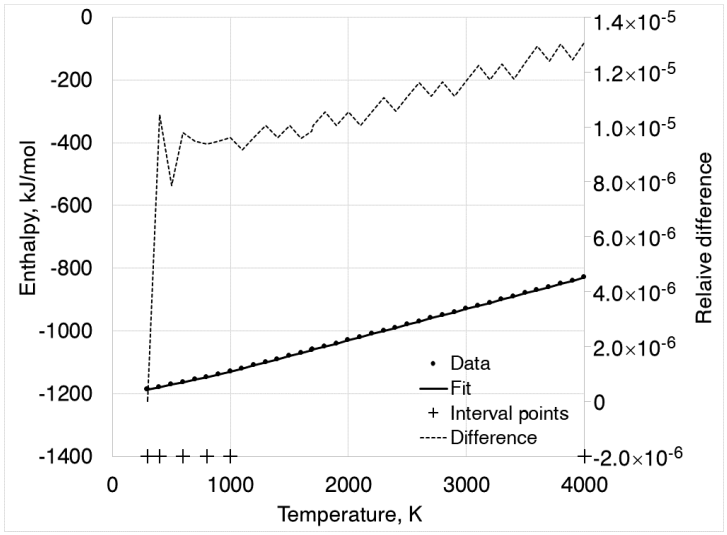


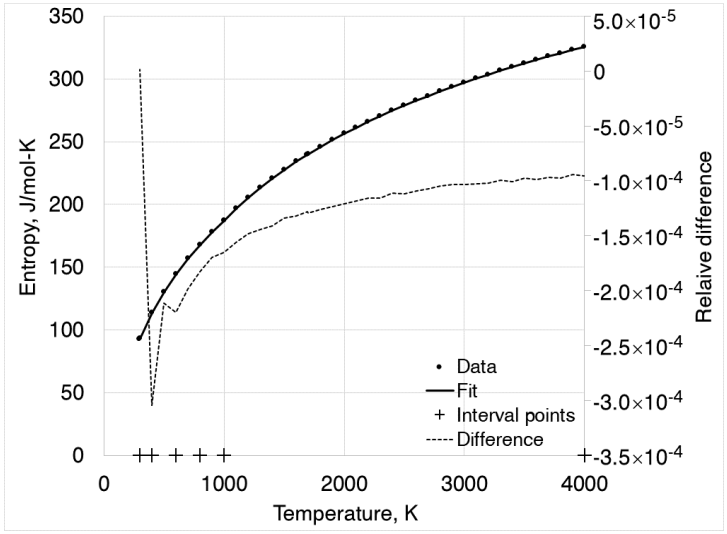


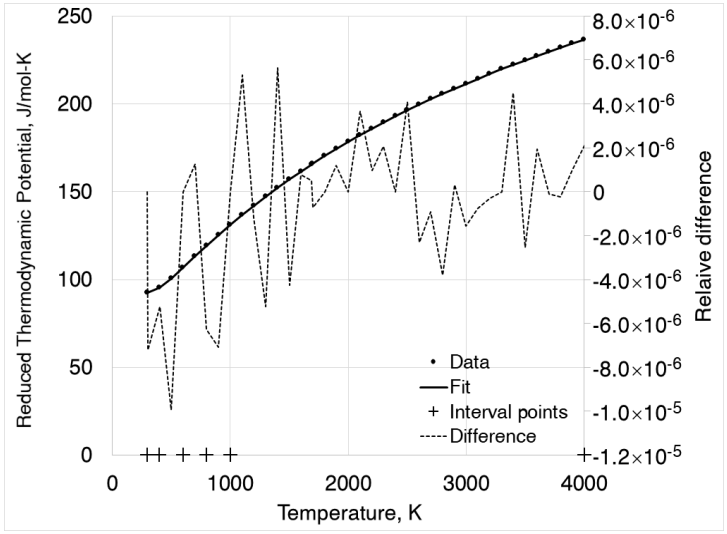
Calcium Fluoride (CaF_2) – Ca1F2(l)

MW = 78.08 g/mol – $\Delta_f H^o$(298.15 K) = -1186.067 kJ/mol – S^o(298.15 K) = 92.573 J/mol-K – p^o = 0.1 MPa



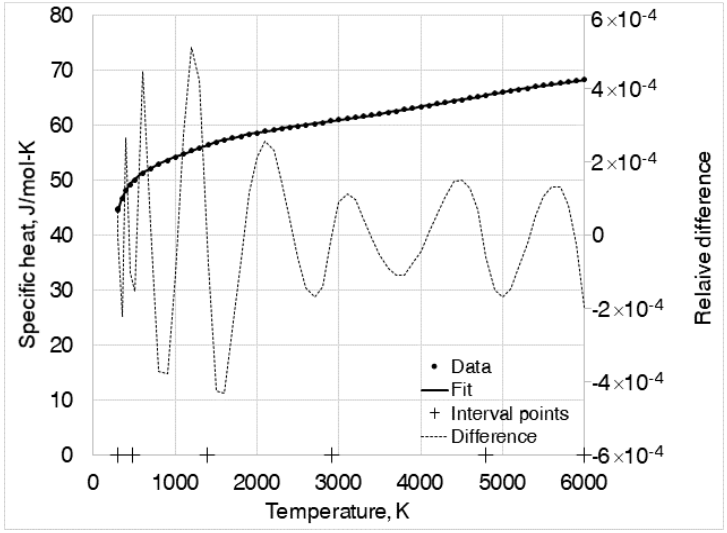


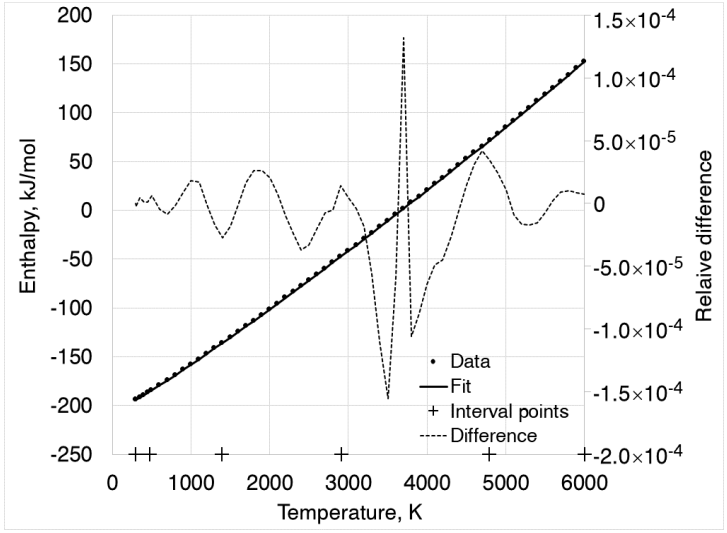


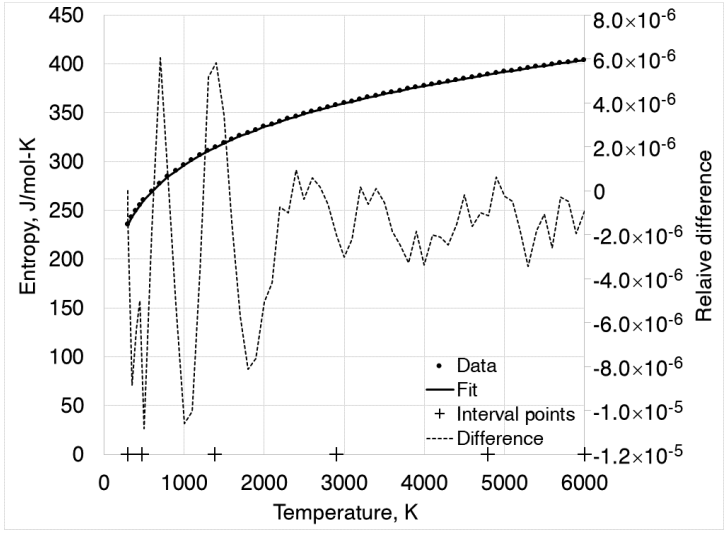


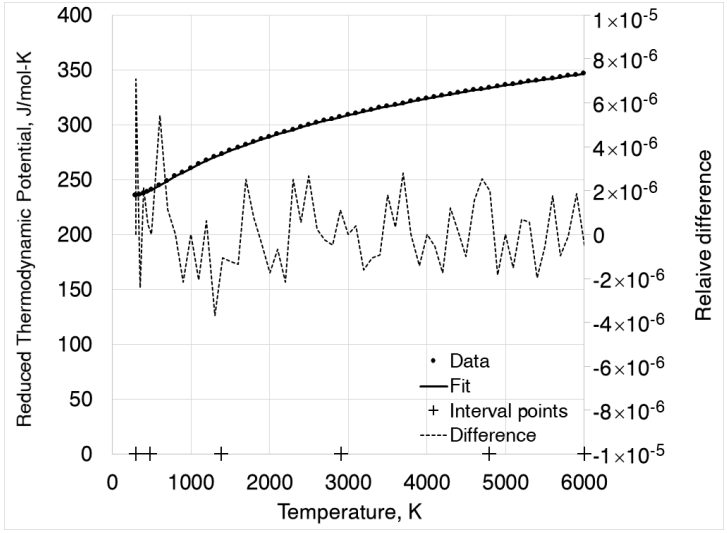
Calcium Hydroxide (CaOH) – Ca1H1O1(g)

MW = 57.09 g/mol – $\Delta_f H^o$(298.15 K) = -193.877 kJ/mol – S^o(298.15 K) = 235.47 J/mol-K – p^o = 0.1 MPa



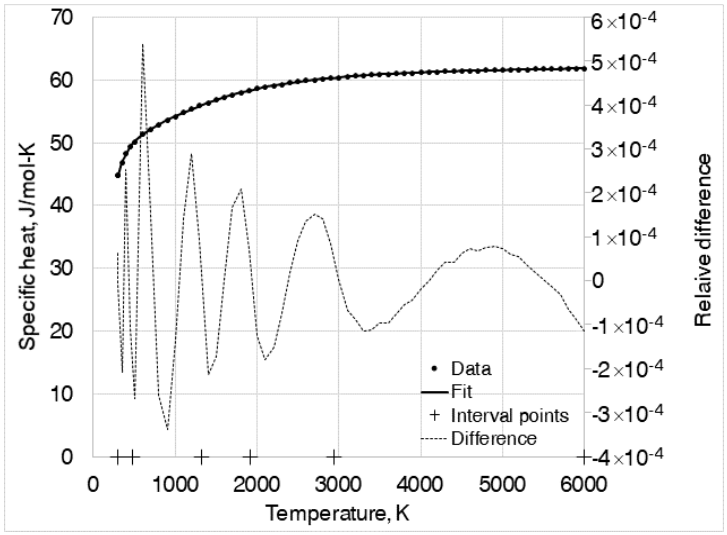


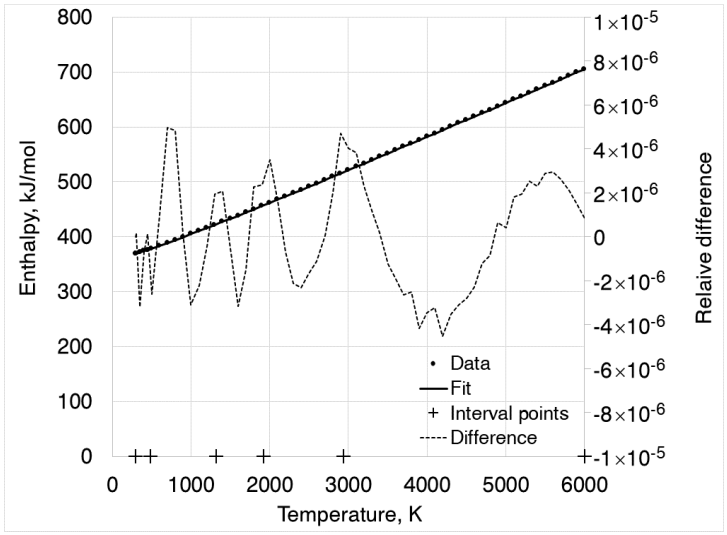


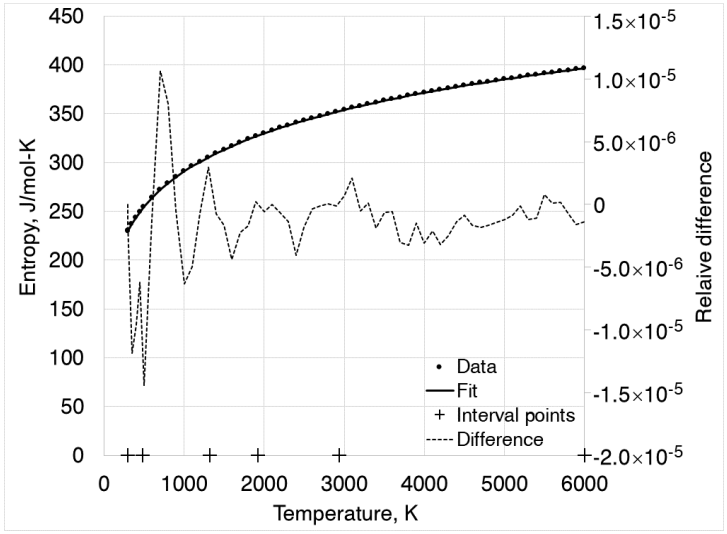


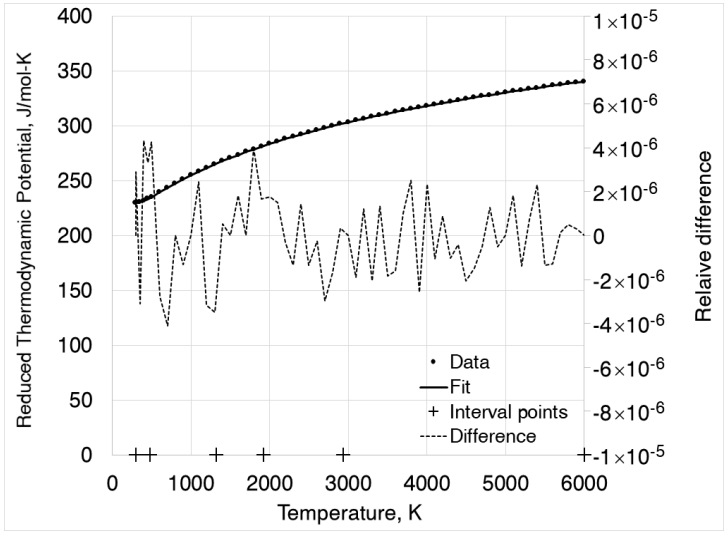
Calcium Hydroxide, Ion ($CaOH^+$) – Ca1H1O1+(g)

MW = 57.09 g/mol – $\Delta_f H^o$(298.15 K) = 369.084 kJ/mol – S^o(298.15 K) = 229.905 J/mol-K – p^o = 0.1 MPa



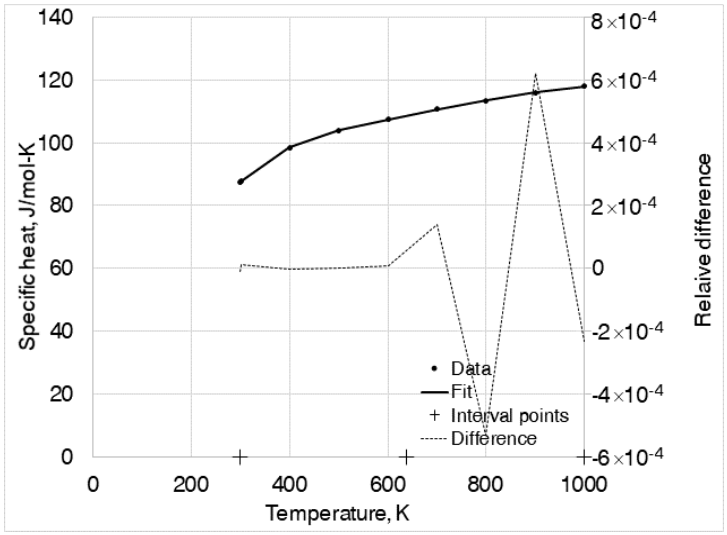


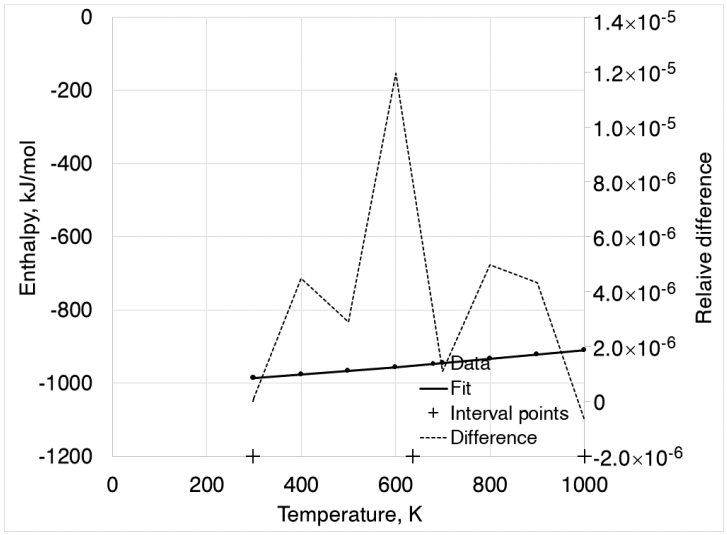


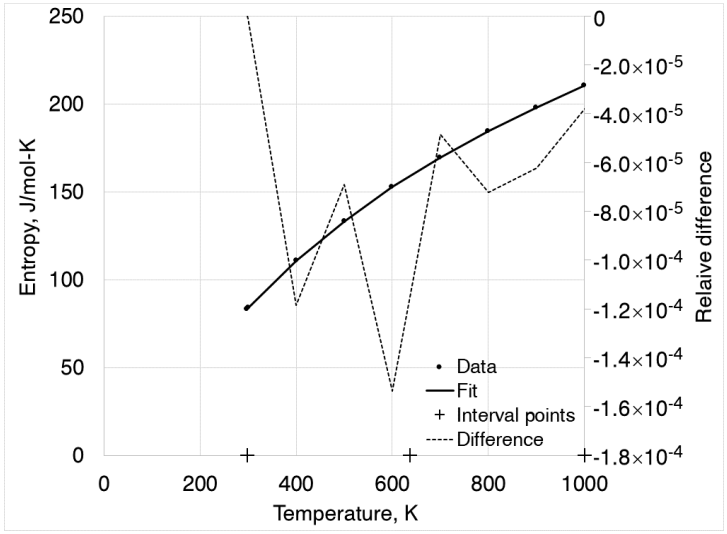


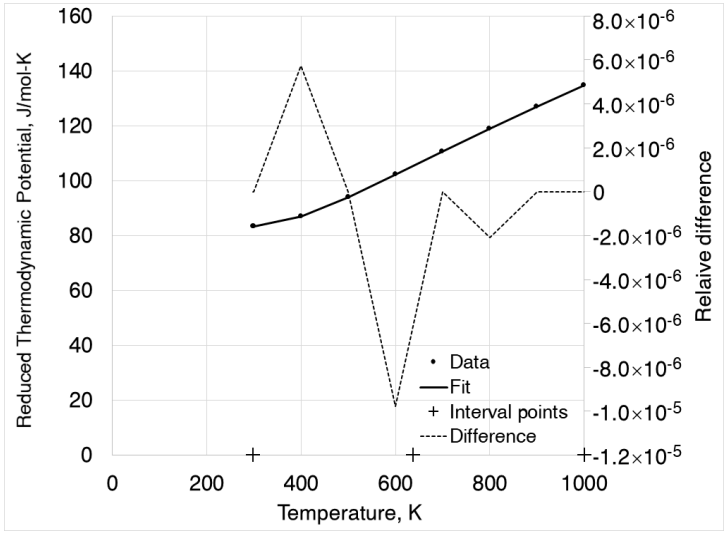
Calcium Hydroxide ($Ca(OH)_2$) – Ca1H2O2(cr)

MW = 74.09 g/mol – $\Delta_f H^o$(298.15 K) = -986.085 kJ/mol – S^o(298.15 K) = 83.387 J/mol-K – p^o = 0.1 MPa



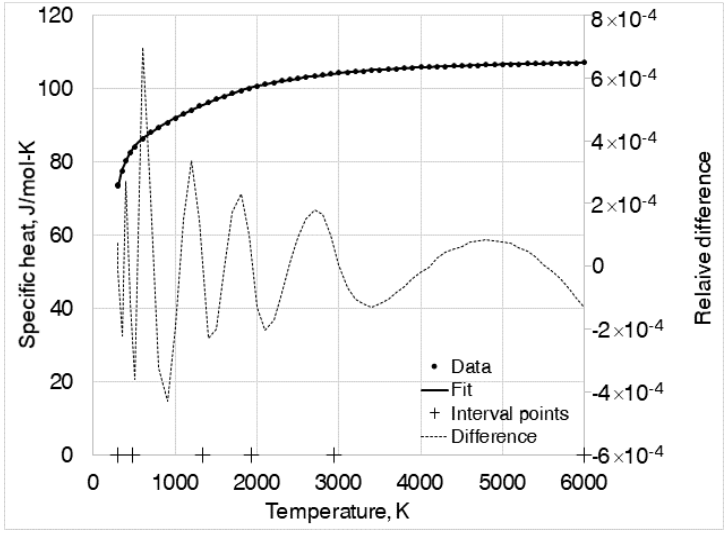


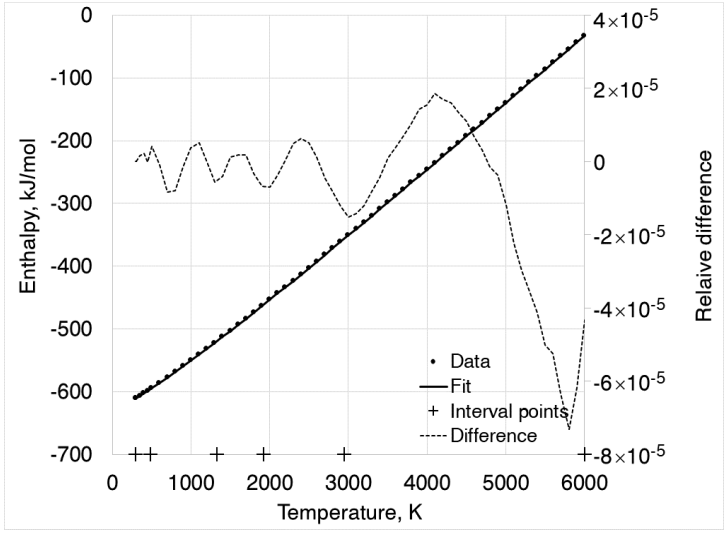


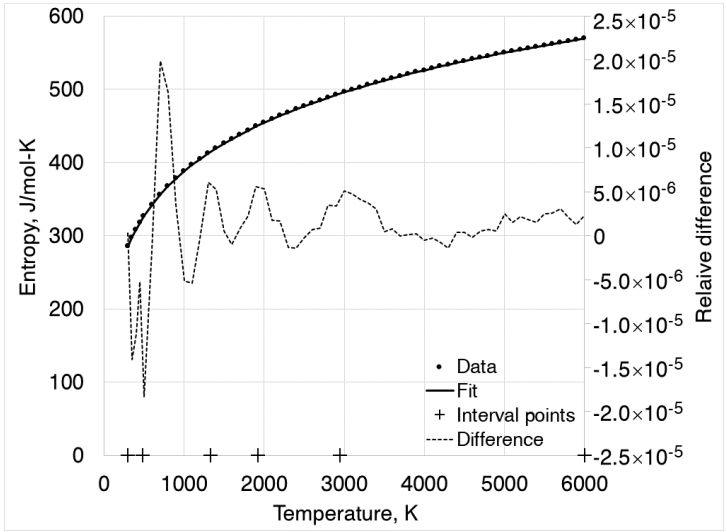


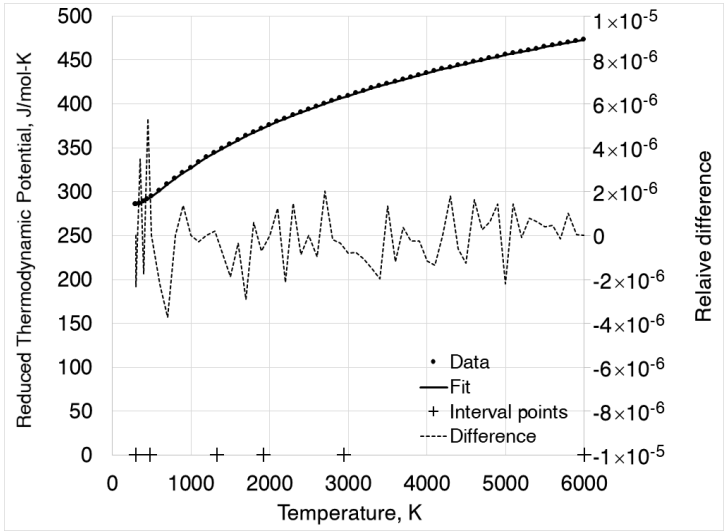
Calcium Hydroxide ($Ca(OH)_2$) – Ca1H2O2(g)

MW = 74.09 g/mol – $\Delta_f H^o$(298.15 K) = -610.764 kJ/mol – S^o(298.15 K) = 285.605 J/mol-K – p^o = 0.1 MPa



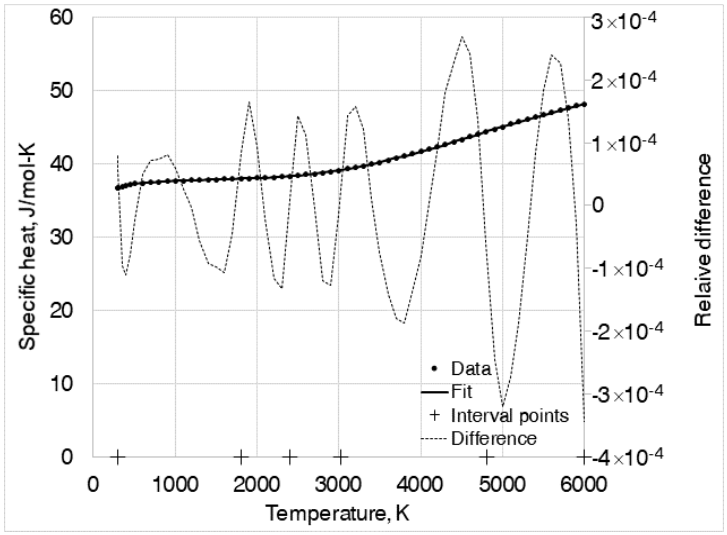


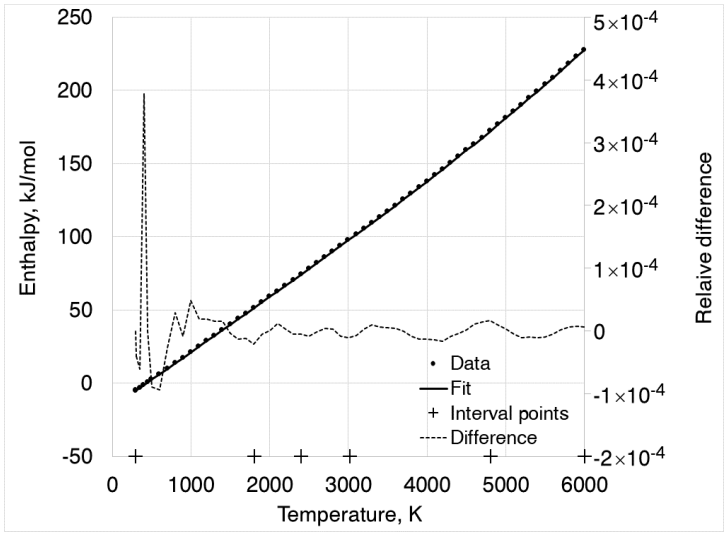


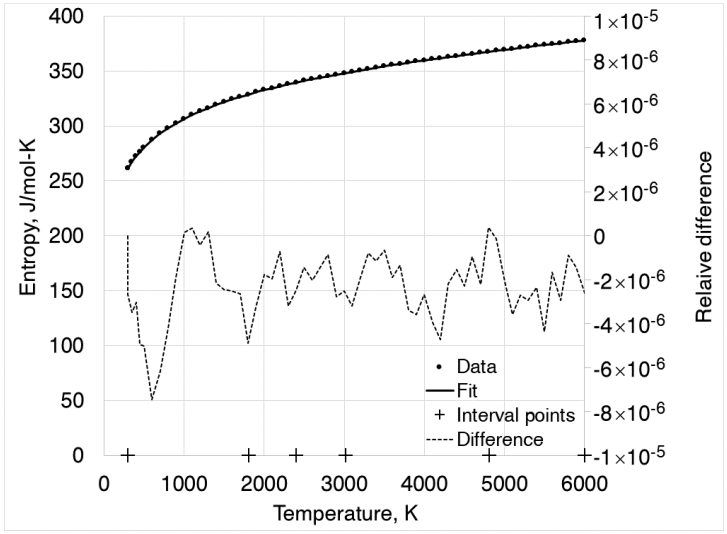


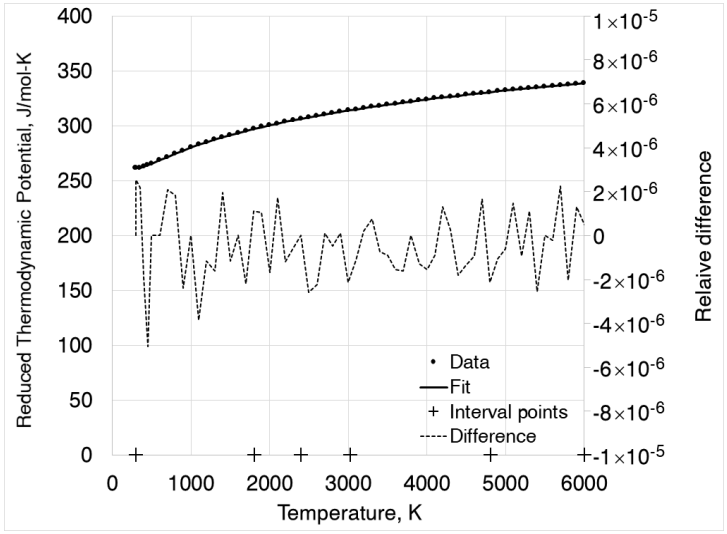
Calcium Iodide (CaI) – Ca1I1(g)

MW = 166.98 g/mol – $\Delta_f H^o$(298.15 K) = -5.045 kJ/mol – S^o(298.15 K) = 261.306 J/mol-K – p^o = 0.1 MPa



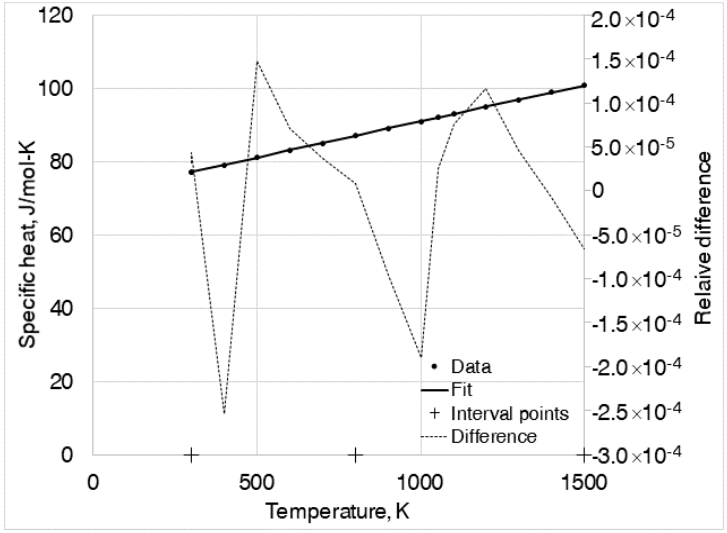


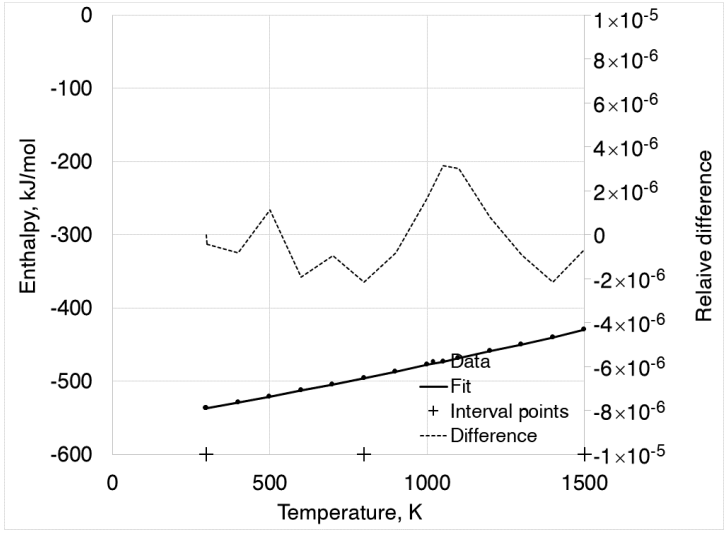


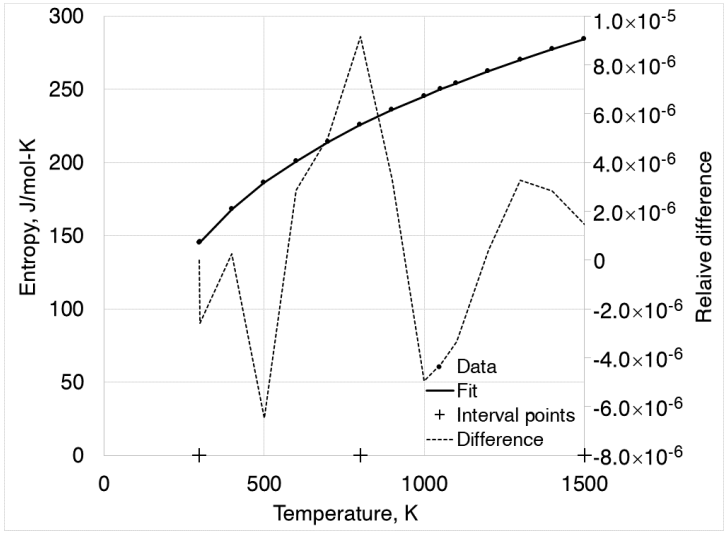


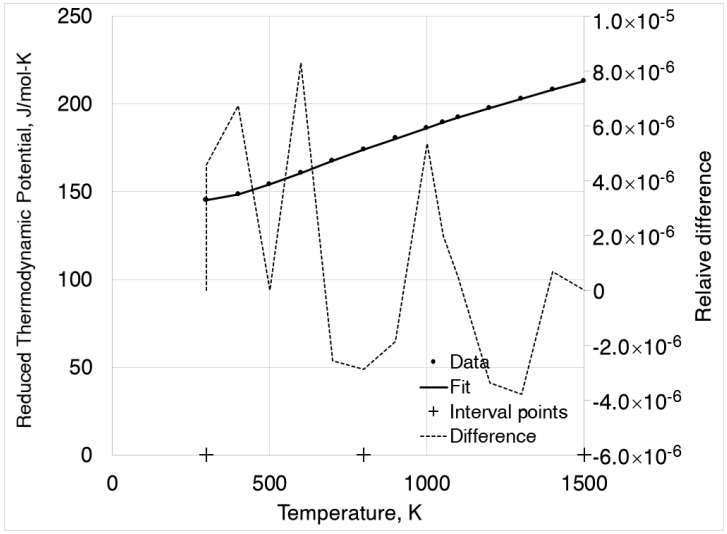
Calcium Iodide (CaI_2) – Ca1I2(cr)

MW = 293.89 g/mol – $\Delta_f H^o$(298.15 K) = -536.807 kJ/mol – S^o(298.15 K) = 145.287 J/mol-K – p^o = 0.1 MPa



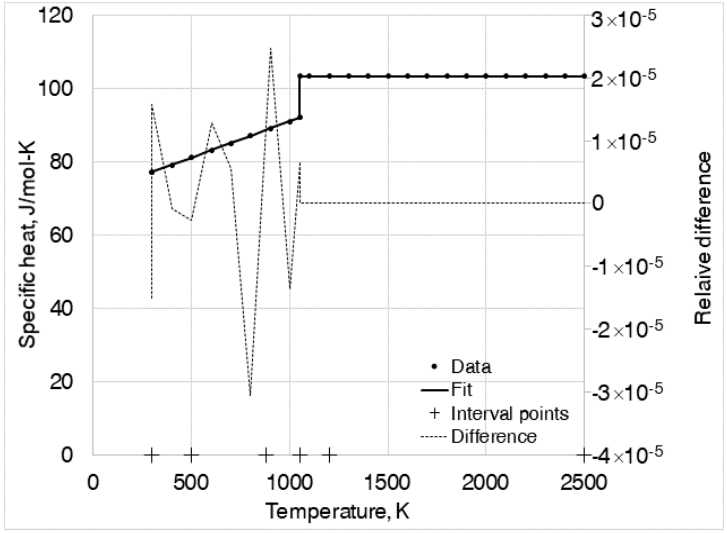


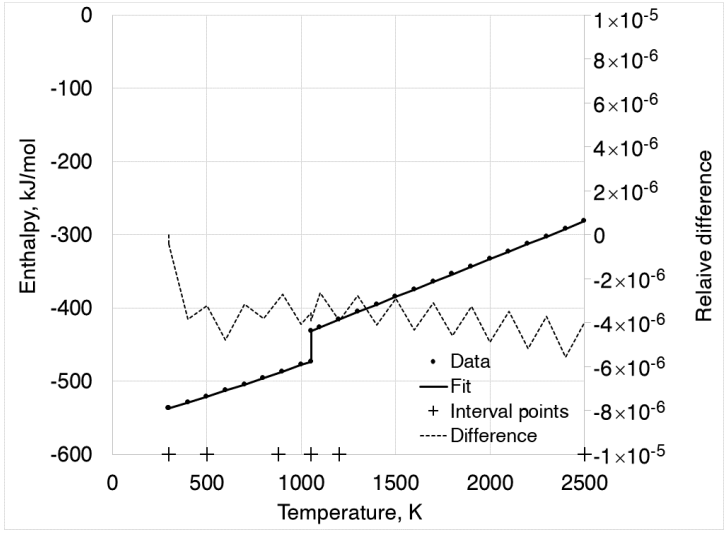


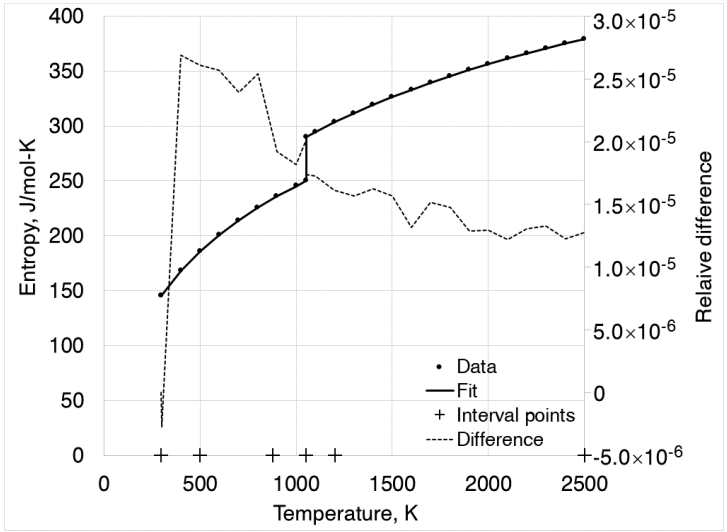


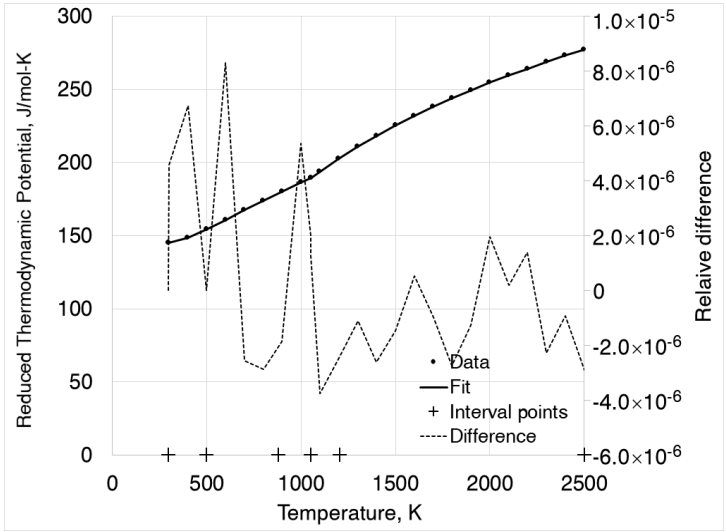
Calcium Iodide (CaI_2) – Ca1I2(cr,l)

MW = 293.89 g/mol – $\Delta_f H^o$(298.15 K) = -536.807 kJ/mol – S^o(298.15 K) = 145.287 J/mol-K – p^o = 0.1 MPa



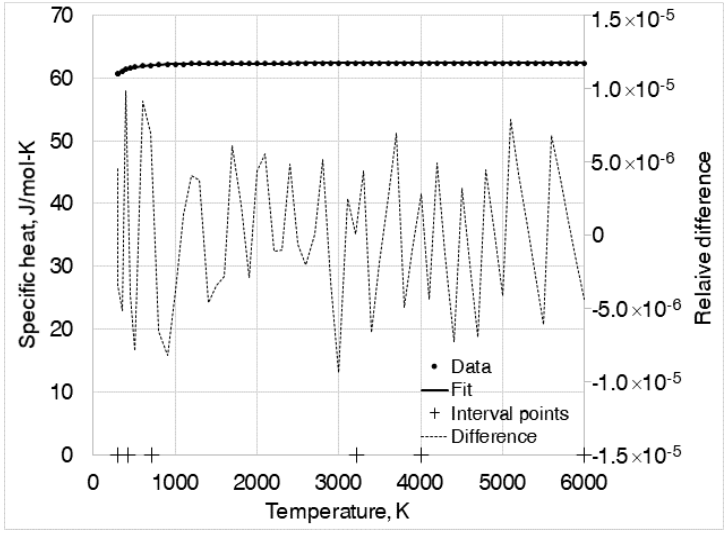


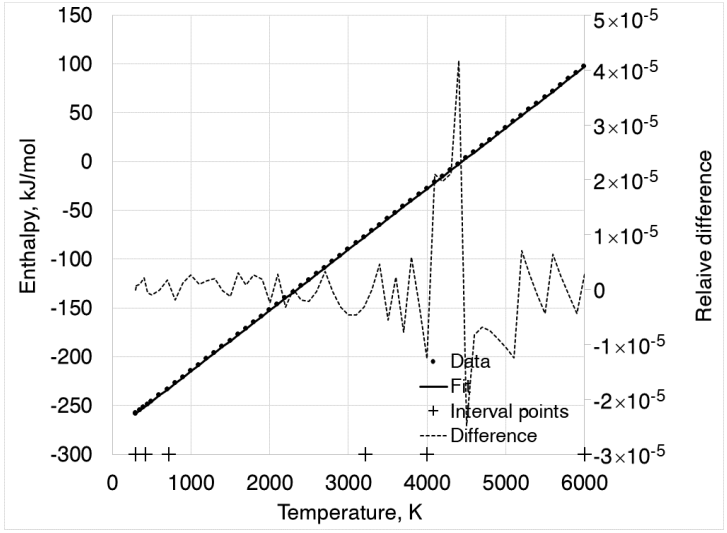


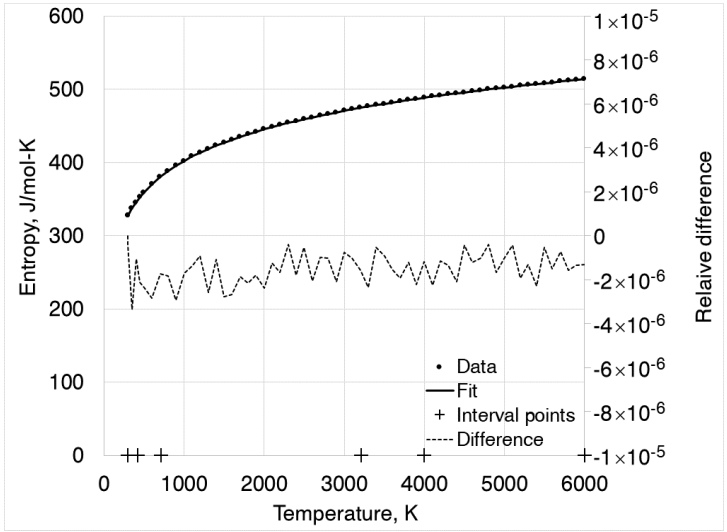


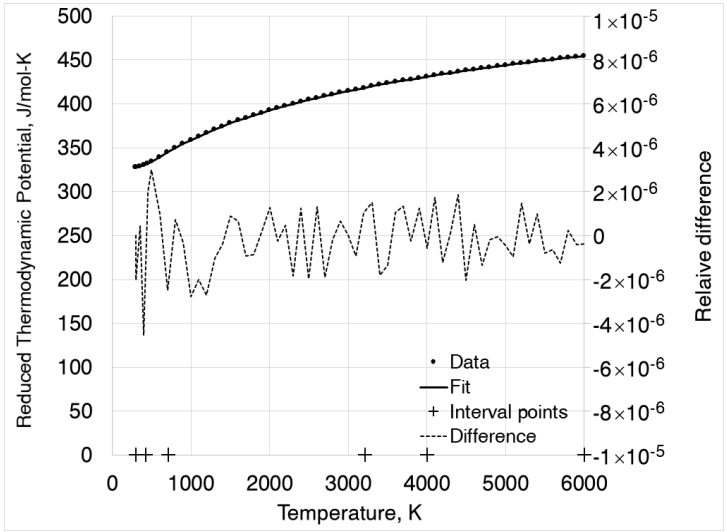
Calcium Iodide (CaI_2) – Ca1I2(g)

MW = 293.89 g/mol – $\Delta_f H^o$(298.15 K) = -258.153 kJ/mol – S^o(298.15 K) = 327.569 J/mol-K – p^o = 0.1 MPa



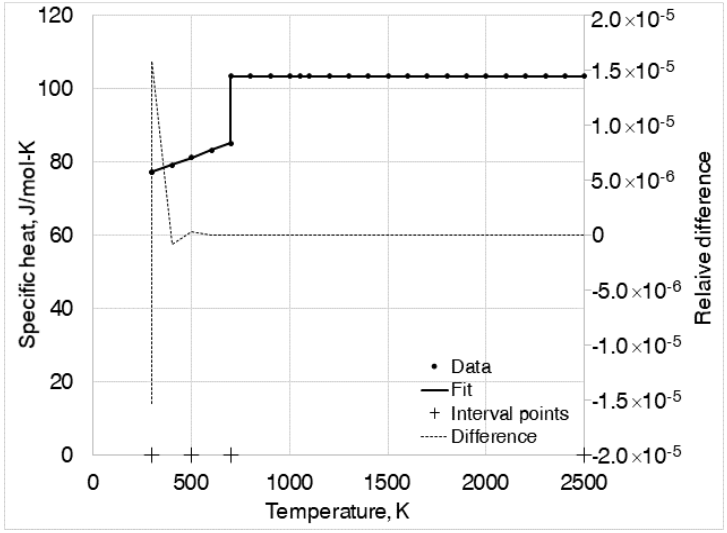


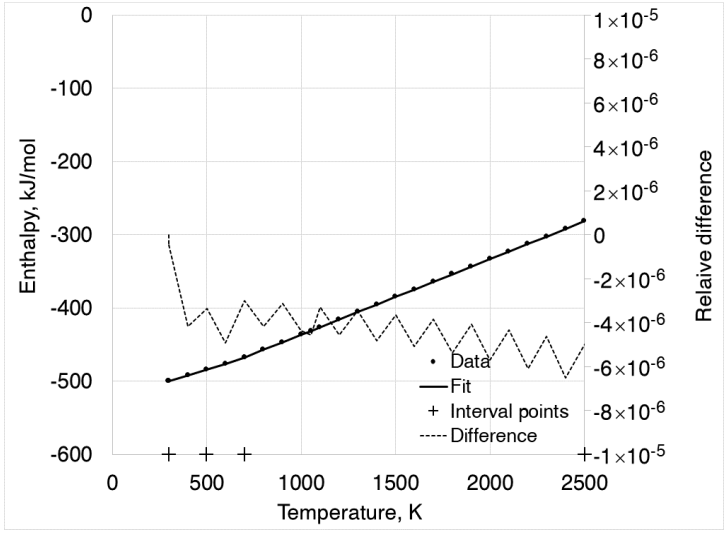


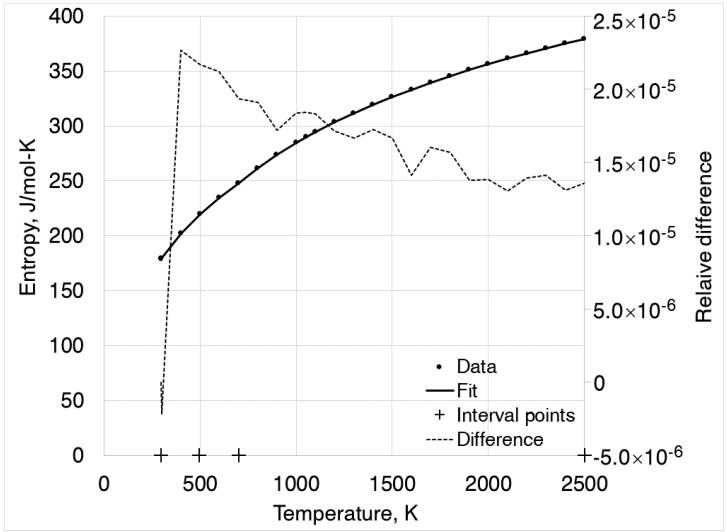


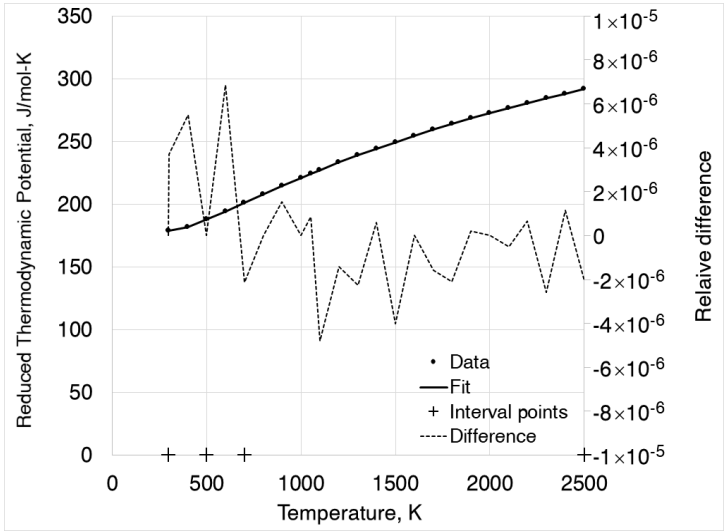
Calcium Iodide (CaI_2) – Ca1I2(l)

MW = 293.89 g/mol – $\Delta_f H^o$(298.15 K) = -500.17 kJ/mol – S^o(298.15 K) = 178.942 J/mol-K – p^o = 0.1 MPa



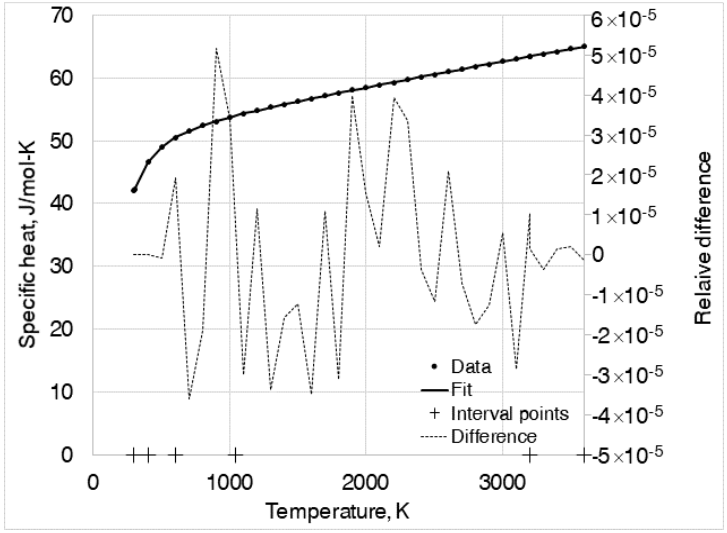


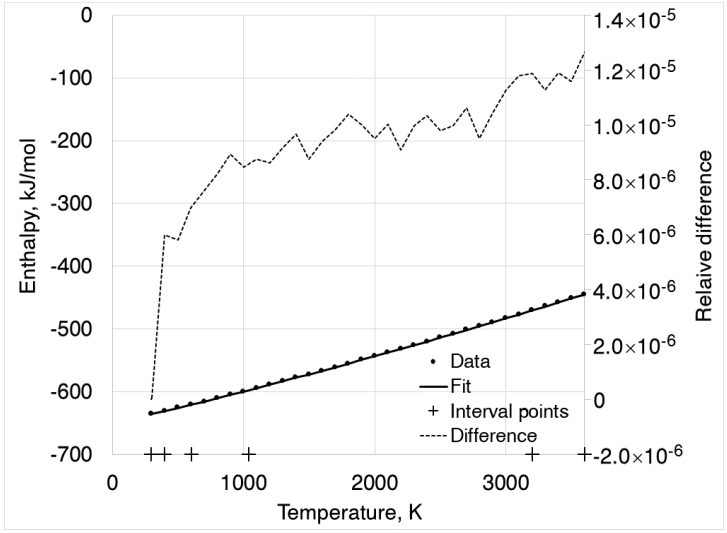


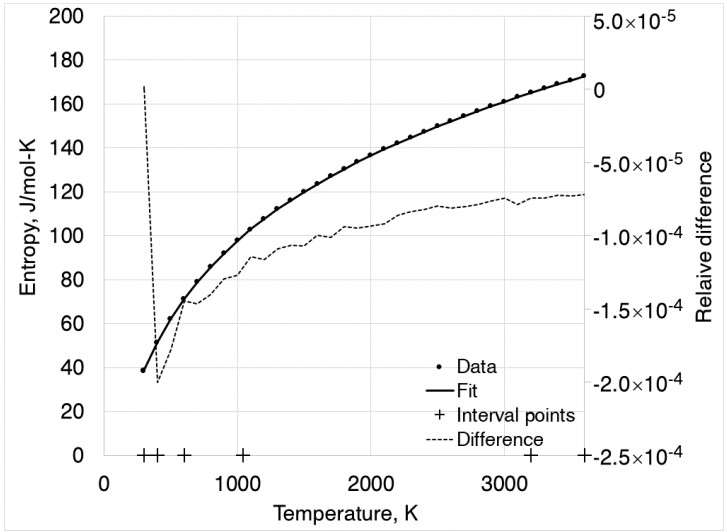


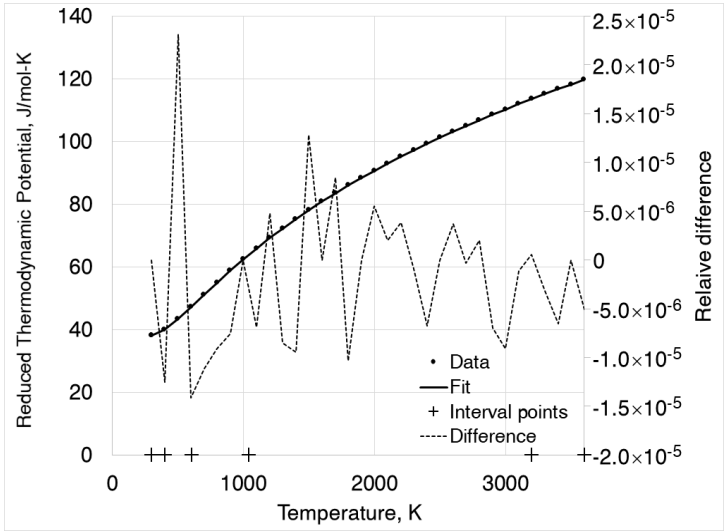
Calcium Oxide (CaO) – Ca1O1(cr)

MW = 56.08 g/mol – $\Delta_f H^o$(298.15 K) = -635.089 kJ/mol – S^o(298.15 K) = 38.212 J/mol-K – p^o = 0.1 MPa



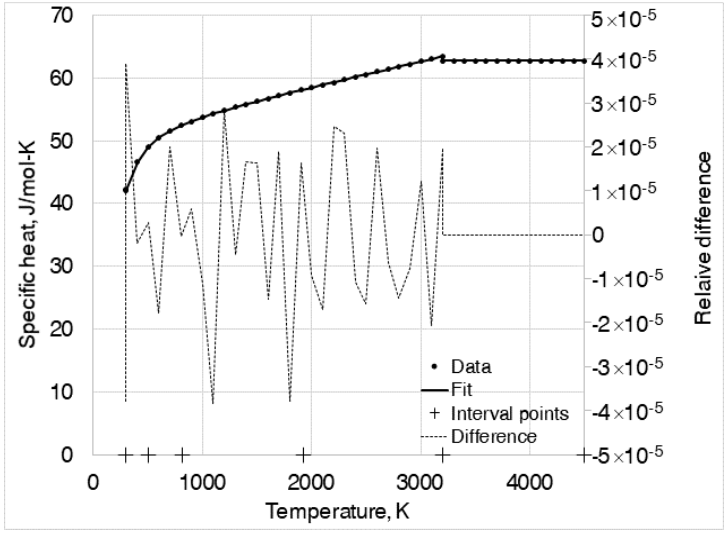


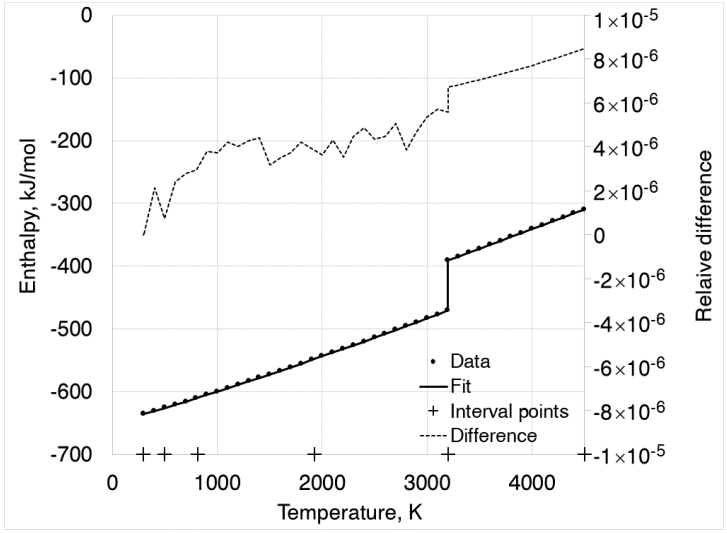


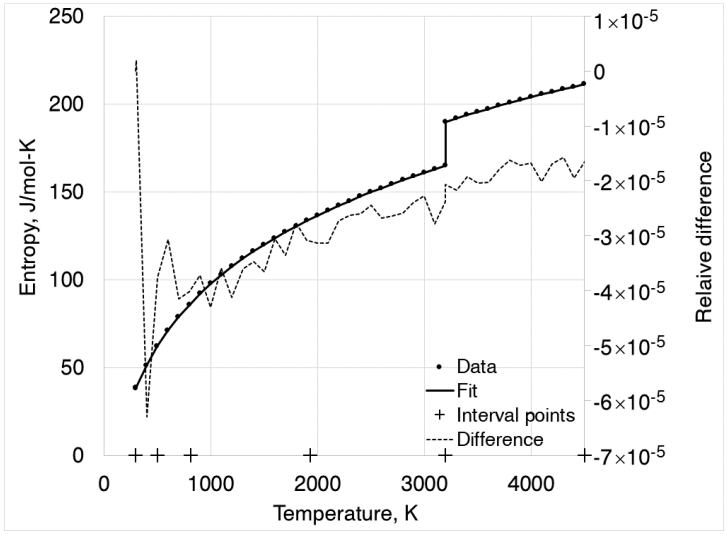


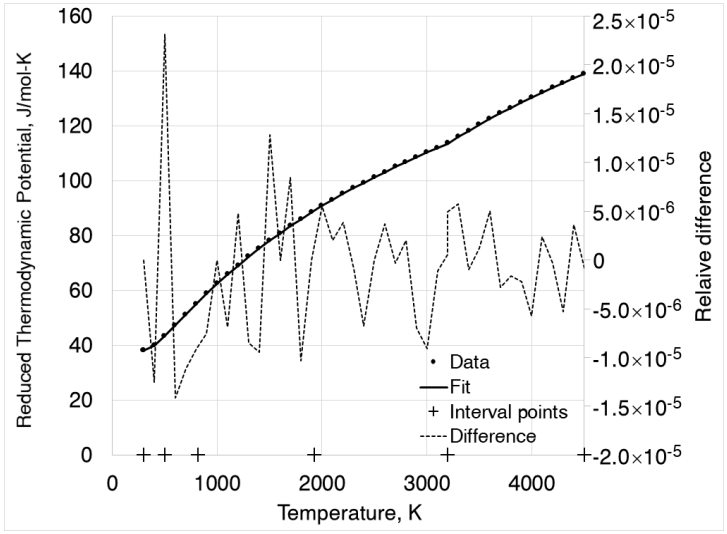
Calcium Oxide (CaO) – Ca1O1(cr,l)

MW = 56.08 g/mol – $\Delta_f H^o$(298.15 K) = -635.089 kJ/mol – S^o(298.15 K) = 38.212 J/mol-K – p^o = 0.1 MPa



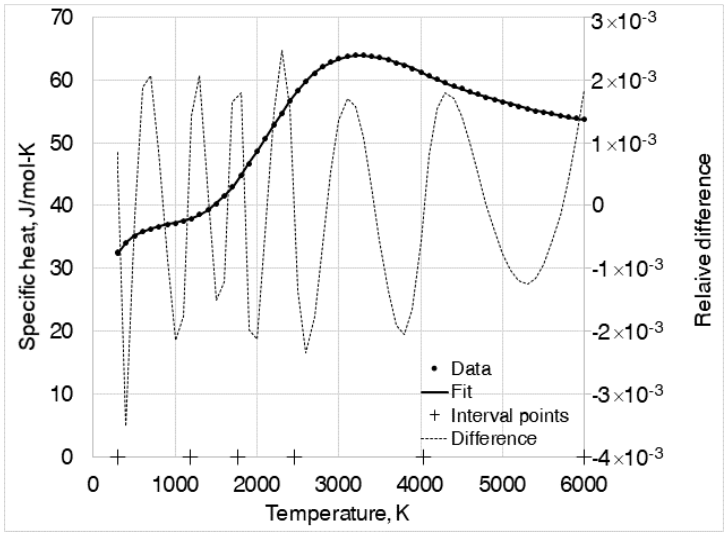


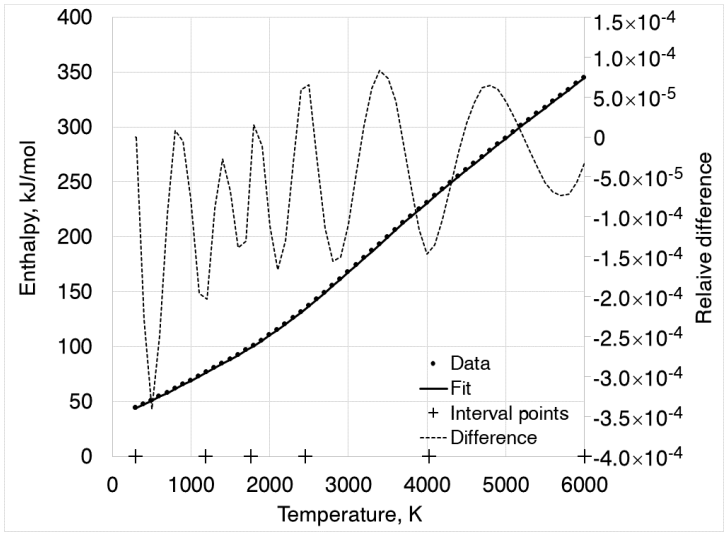


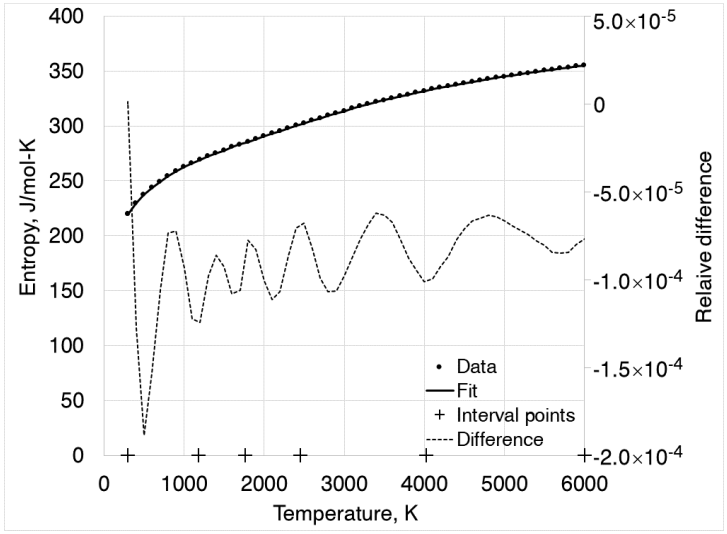


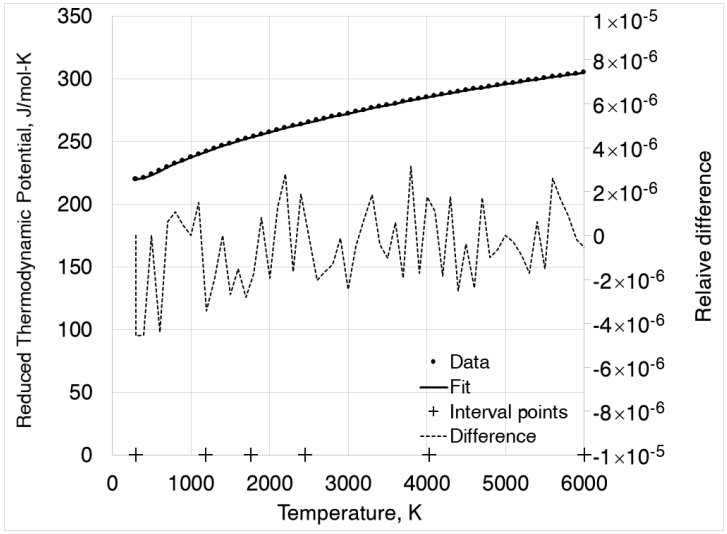
Calcium Oxide (CaO) – Ca1O1(g)

MW = 56.08 g/mol – $\Delta_f H^o$(298.15 K) = 43.932 kJ/mol – S^o(298.15 K) = 219.717 J/mol-K – p^o = 0.1 MPa



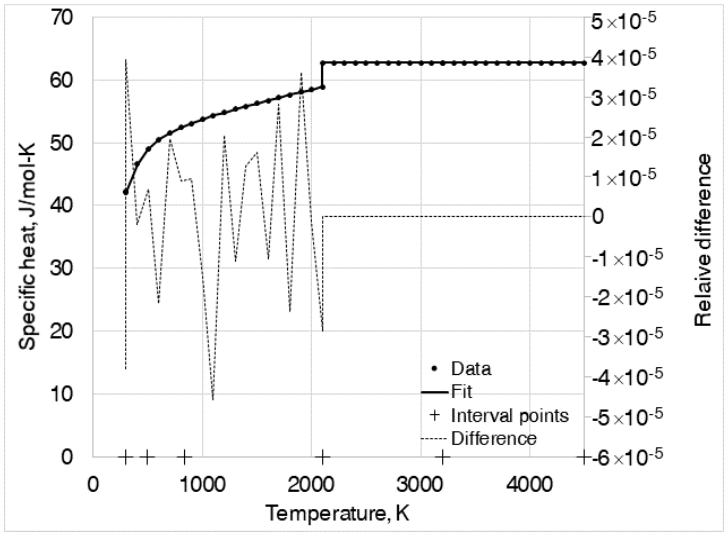


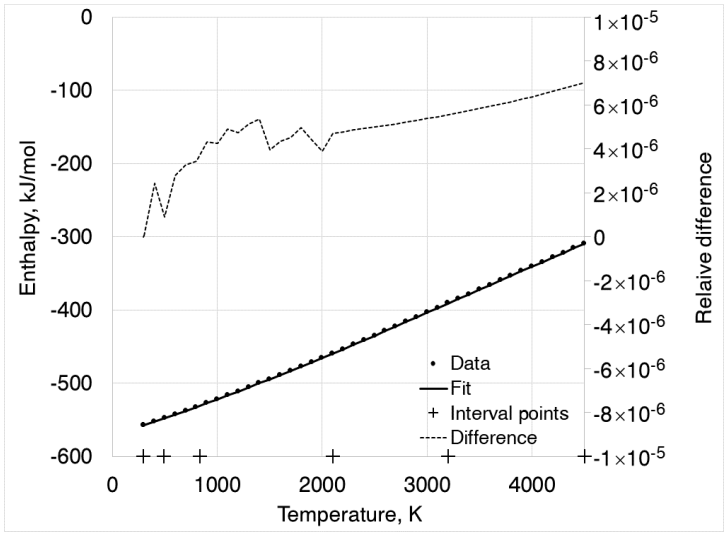


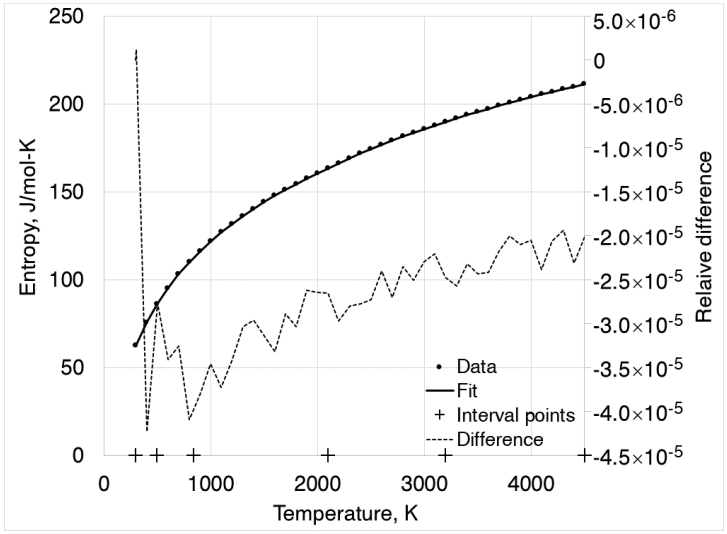


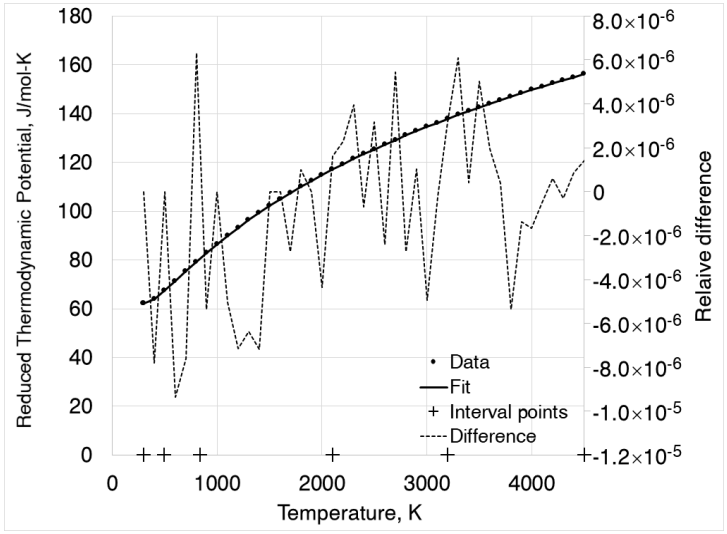
Calcium Oxide (CaO) – Ca1O1(l)

MW = 56.08 g/mol – $\Delta_f H^o$(298.15 K) = -557.335 kJ/mol – S^o(298.15 K) = 62.321 J/mol-K – p^o = 0.1 MPa



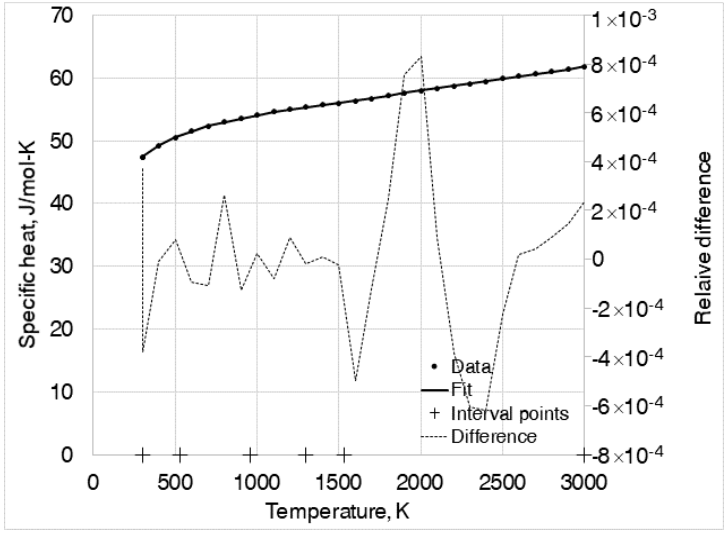


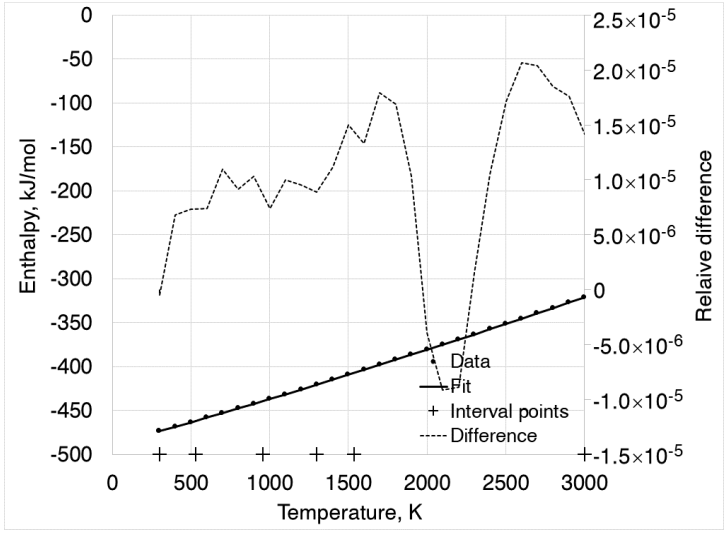


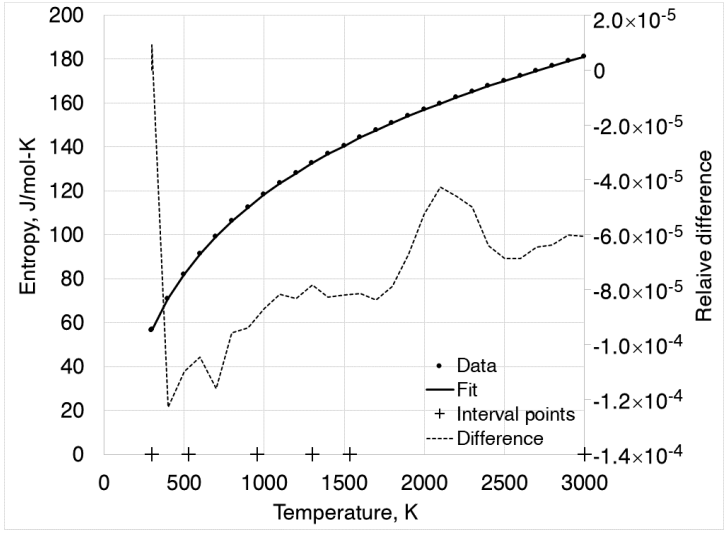


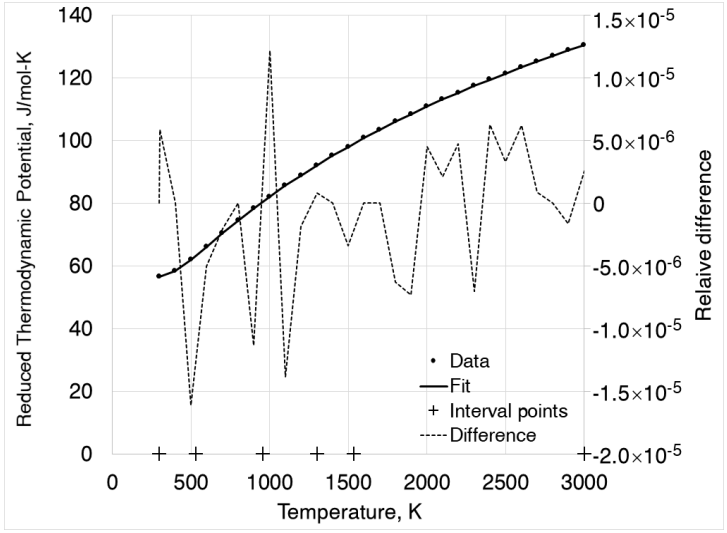
Calcium Sulfide (CaS) – Ca1S1(cr)

MW = 72.14 g/mol – $\Delta_f H^o$(298.15 K) = -473.21 kJ/mol – S^o(298.15 K) = 56.601 J/mol-K – p^o = 0.1 MPa



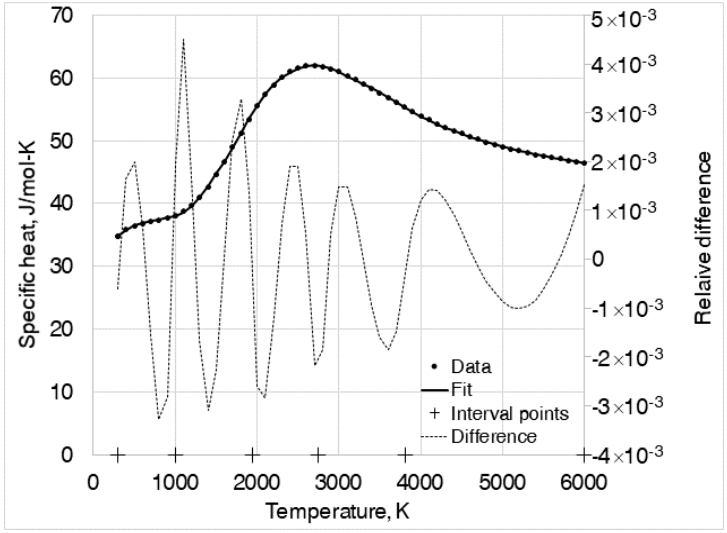


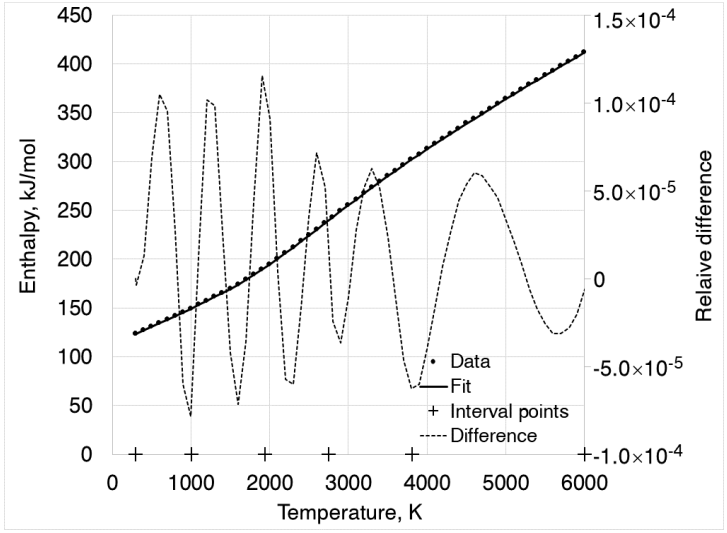


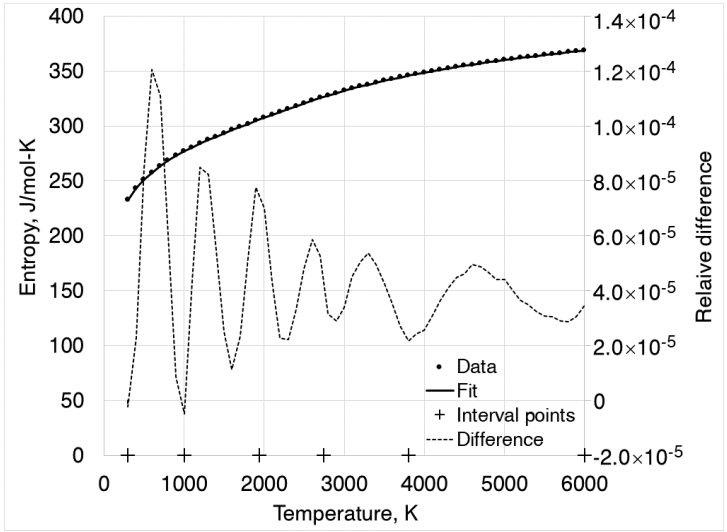


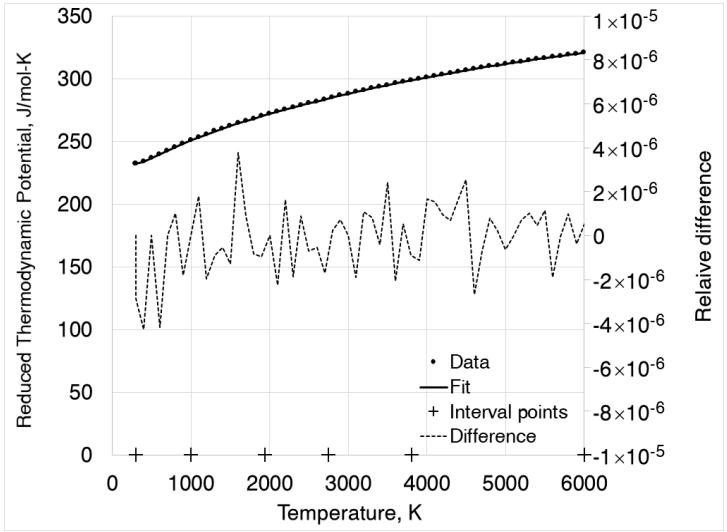
Calcium Sulfide (CaS) – Ca1S1(g)

MW = 72.14 g/mol – $\Delta_f H^o$(298.15 K) = 123.595 kJ/mol – S^o(298.15 K) = 232.589 J/mol-K – p^o = 0.1 MPa



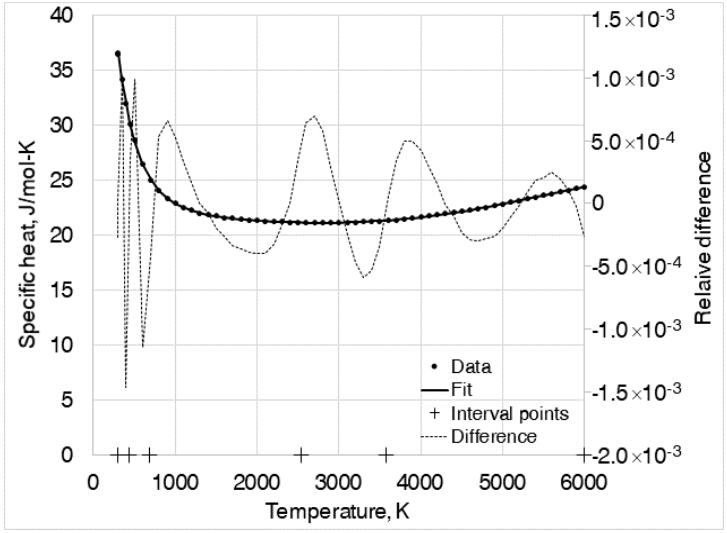


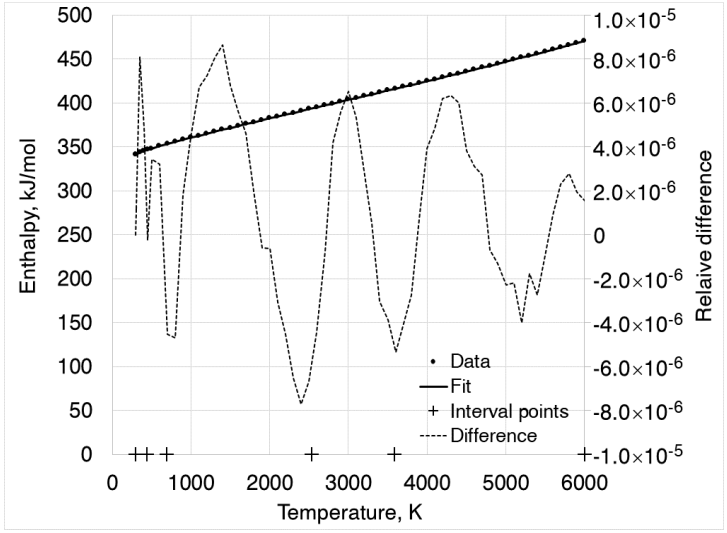


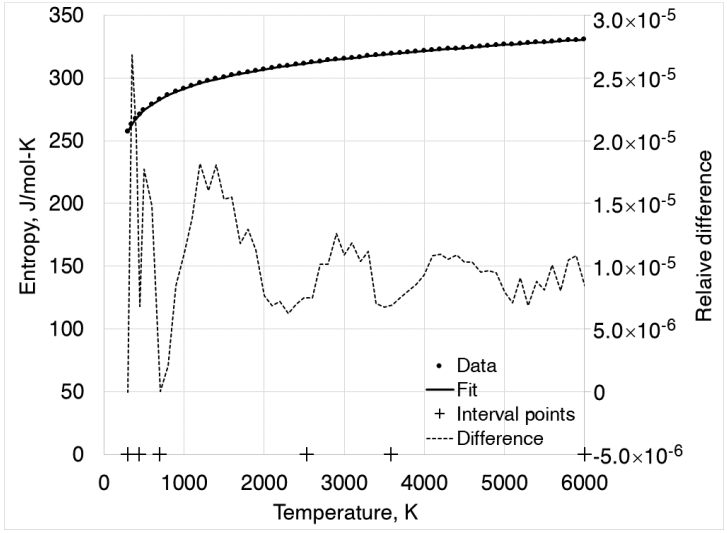


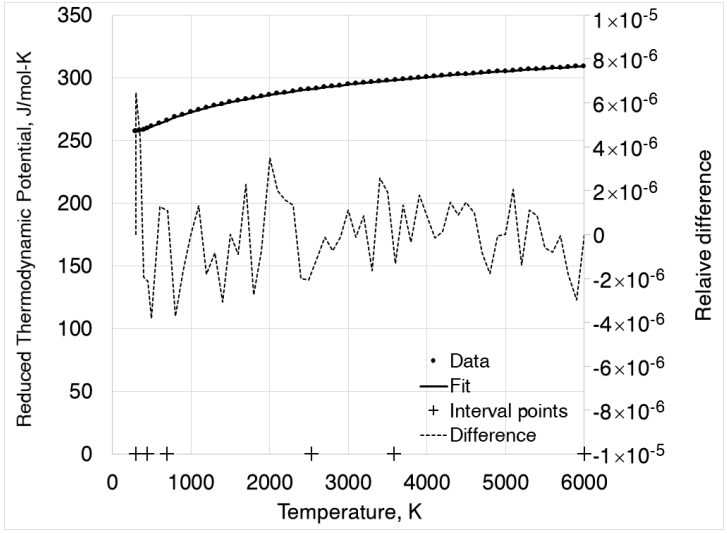
Calcium (Ca_2) – Ca2(g)

MW = 80.16 g/mol – $\Delta_f H^o$(298.15 K) = 341.637 kJ/mol – S^o(298.15 K) = 257.183 J/mol-K – p^o = 0.1 MPa



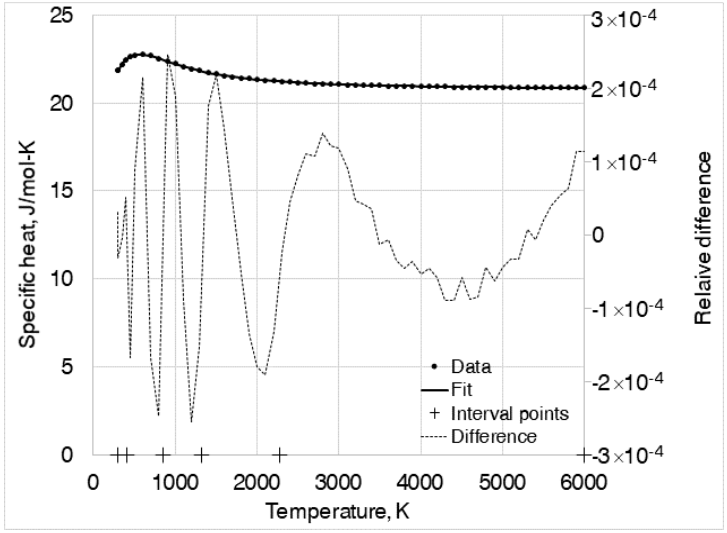


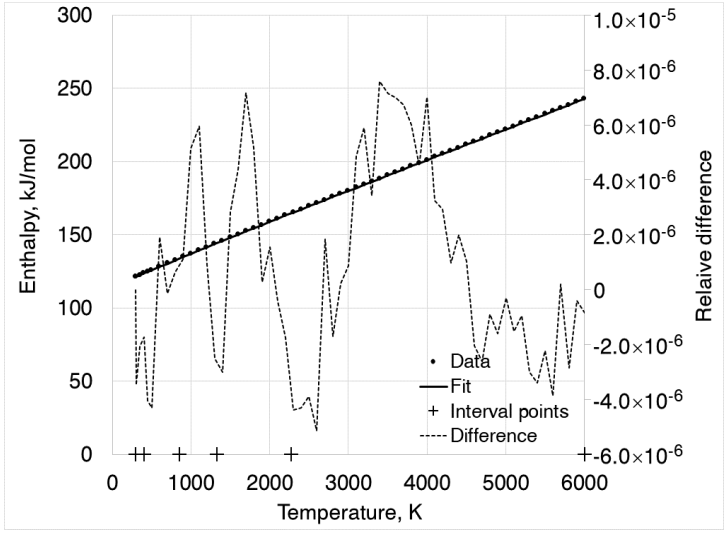


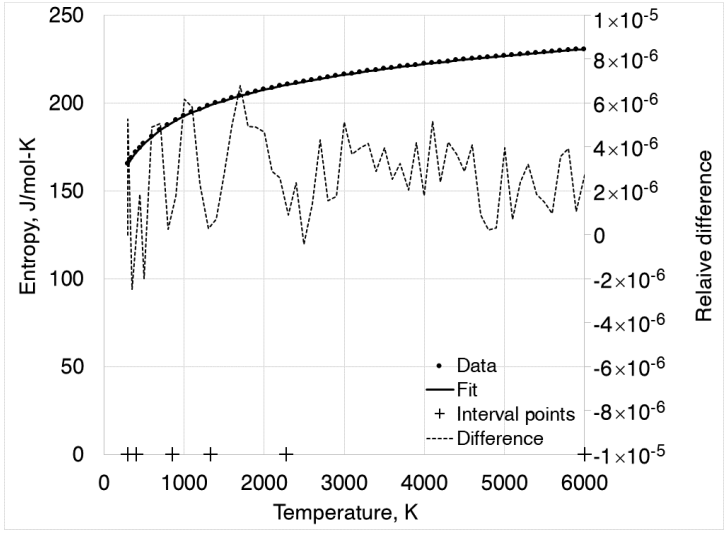


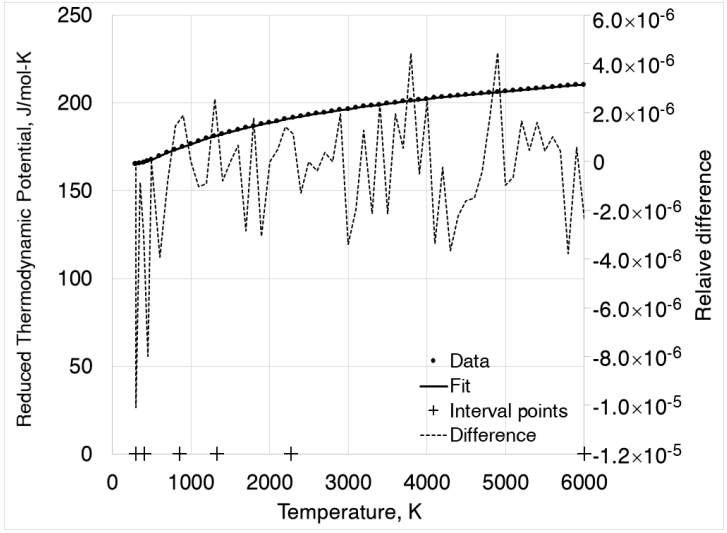
Chlorine (Cl) – Cl1(g)

MW = 35.453 g/mol – $\Delta_f H^o$(298.15 K) = 121.302 kJ/mol – S^o(298.15 K) = 165.189 J/mol-K – p^o = 0.1 MPa



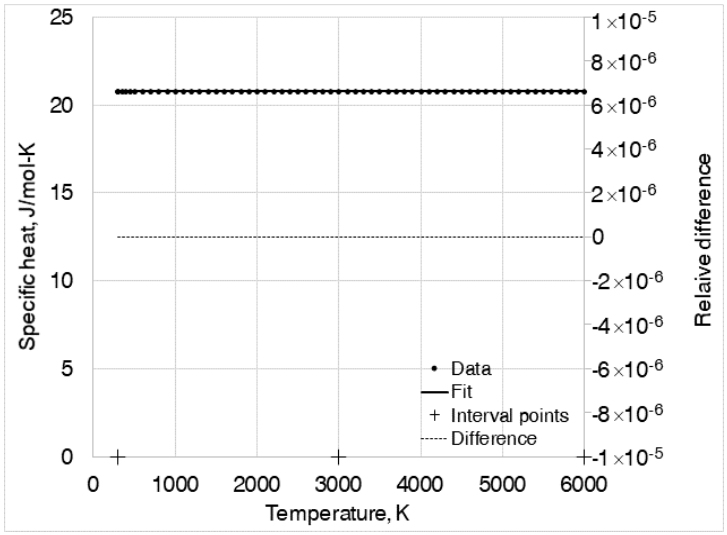


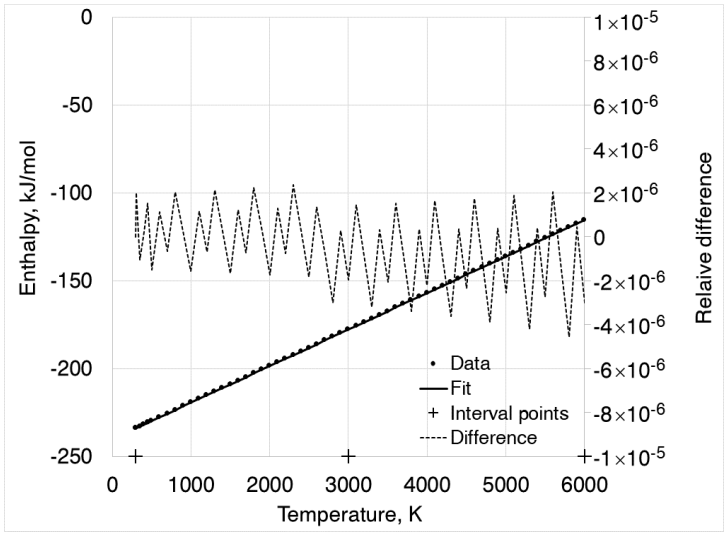


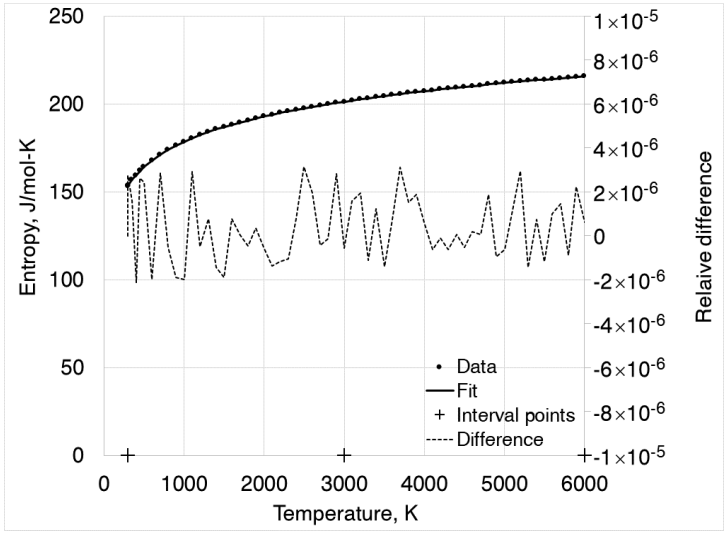


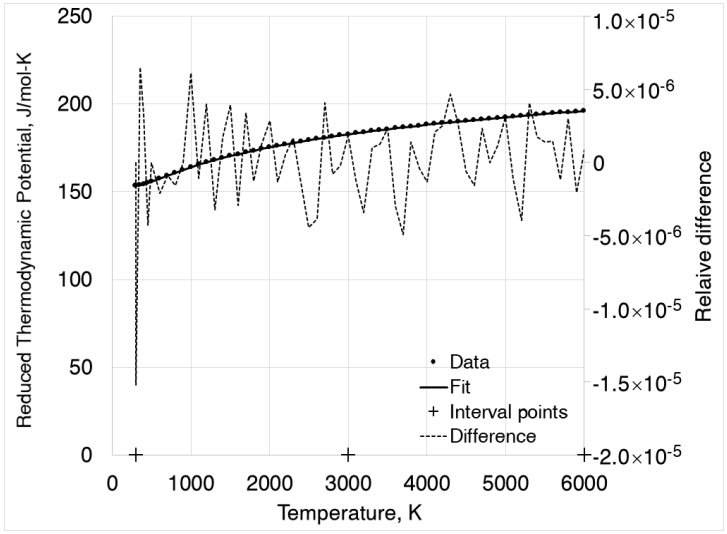
Chlorine, Ion (Cl^-) – Cl1-(g)

MW = 35.454 g/mol – $\Delta_f H^o$(298.15 K) = -233.954 kJ/mol – S^o(298.15 K) = 153.356 J/mol-K – p^o = 0.1 MPa



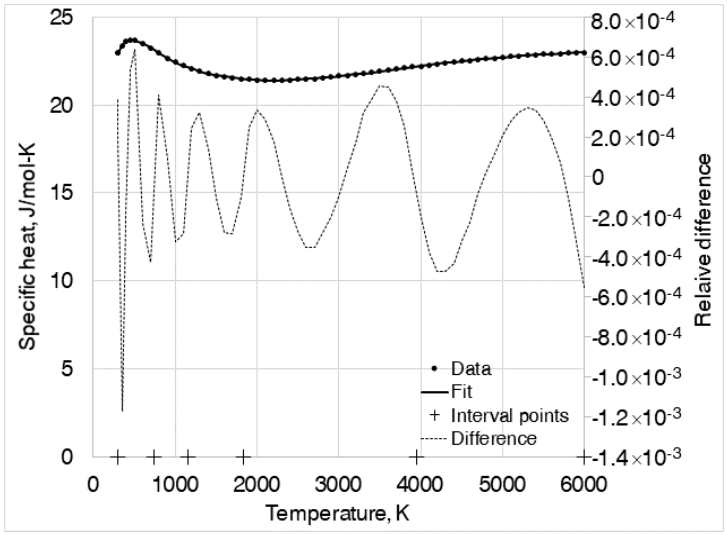


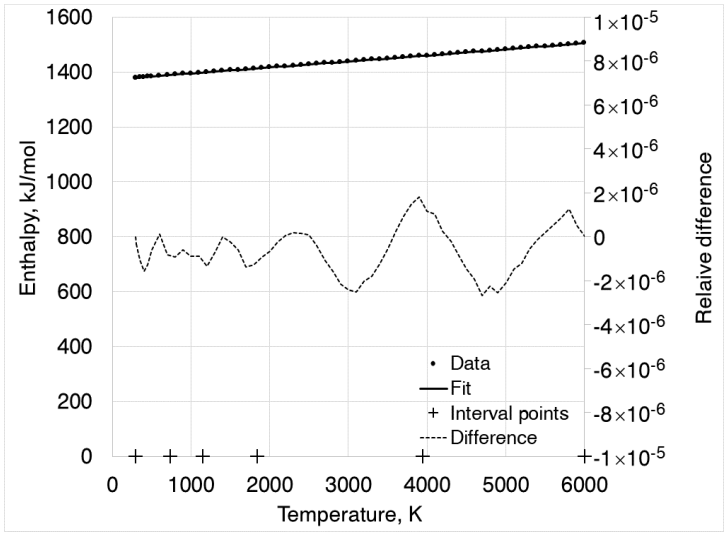


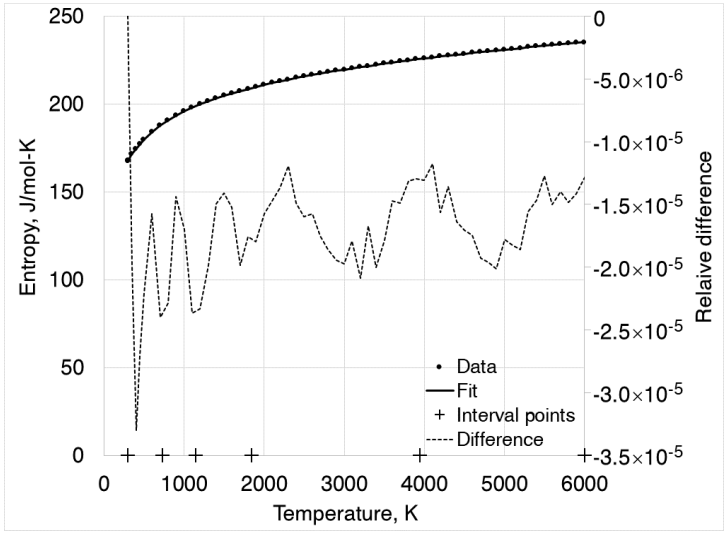


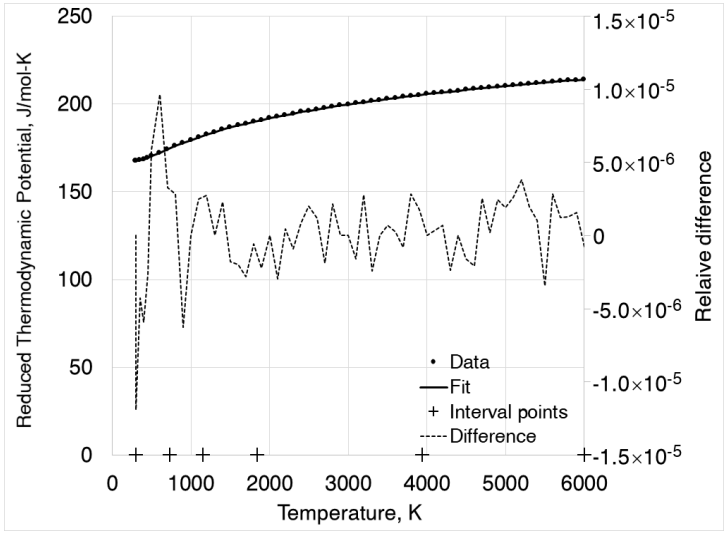
Chlorine, Ion (Cl^+) – Cl1+(g)

MW = 35.452 g/mol – $\Delta_f H^o$(298.15 K) = 1378.801 kJ/mol – S^o(298.15 K) = 167.556 J/mol-K – p^o = 0.1 MPa



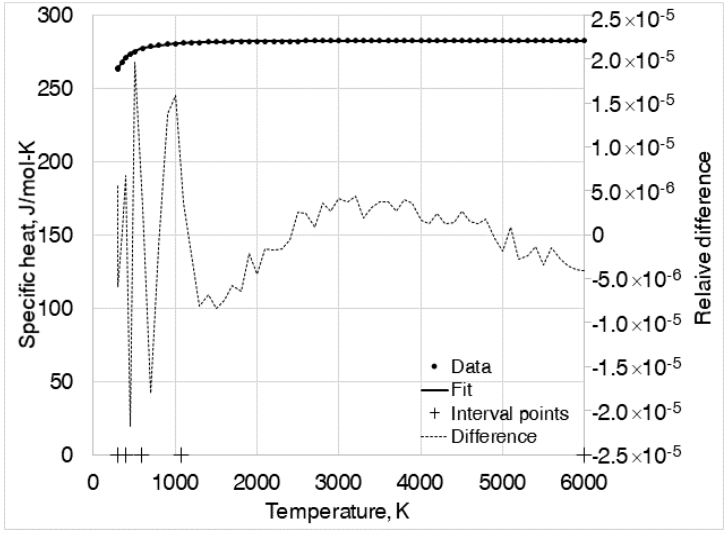


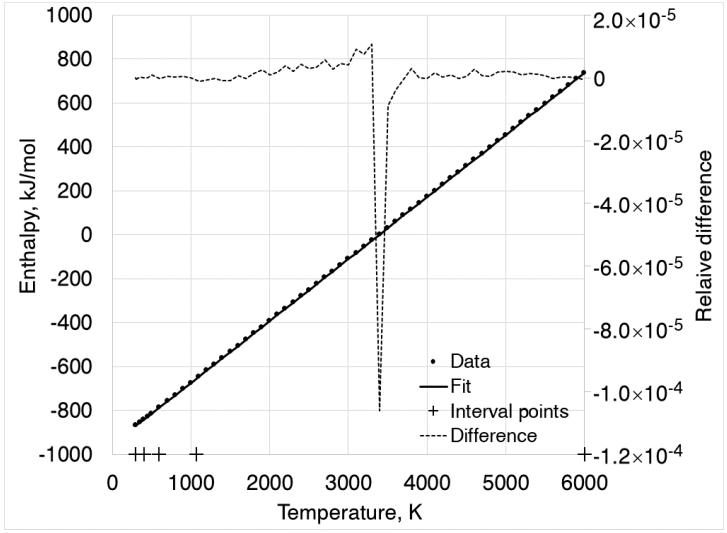


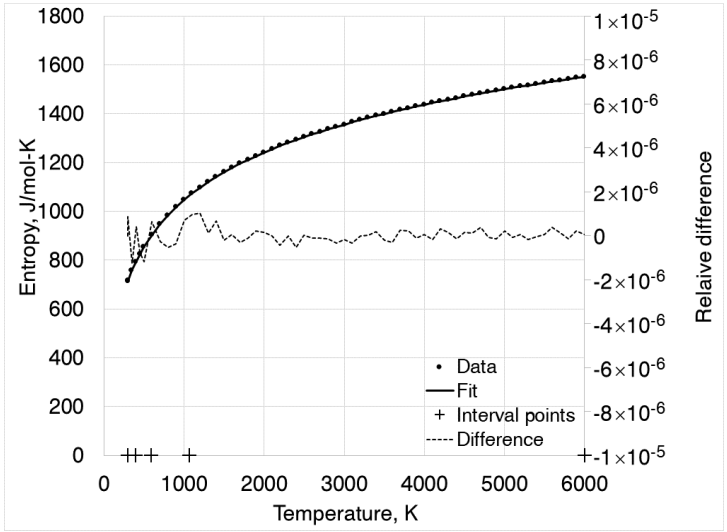


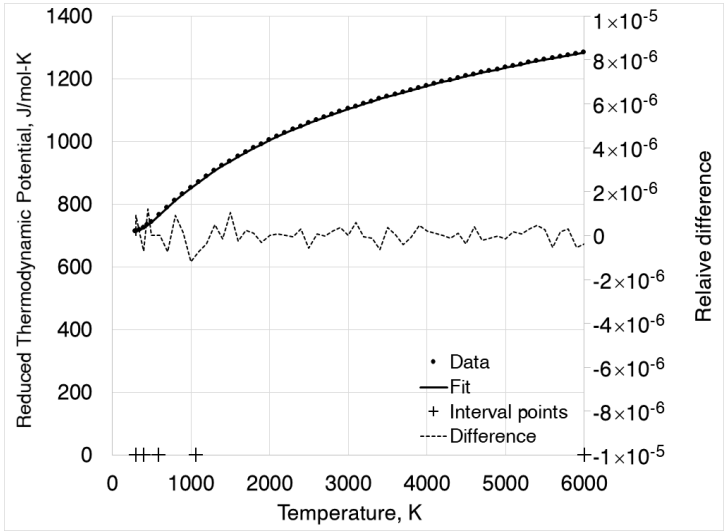
Tungsten Chloride ($(WCl_5)_2$) – Cl10W2(g)

MW = 722.23 g/mol – $\Delta_f H^o$(298.15 K) = -868.598 kJ/mol – S^o(298.15 K) = 713.581 J/mol-K – p^o = 0.1 MPa



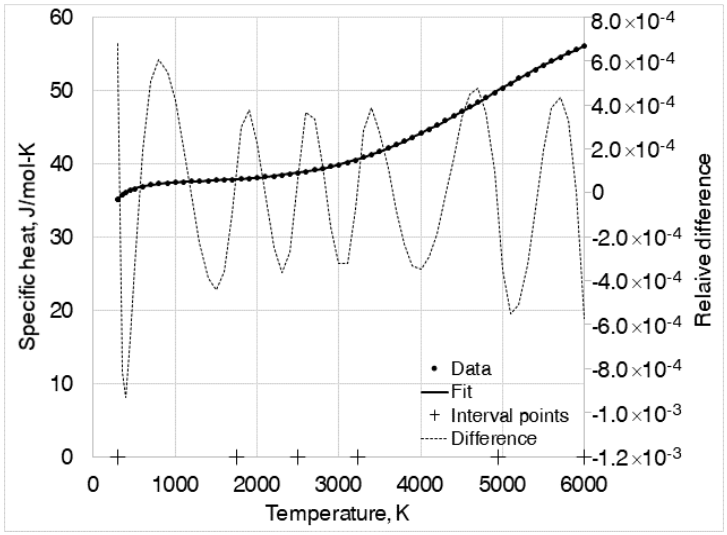


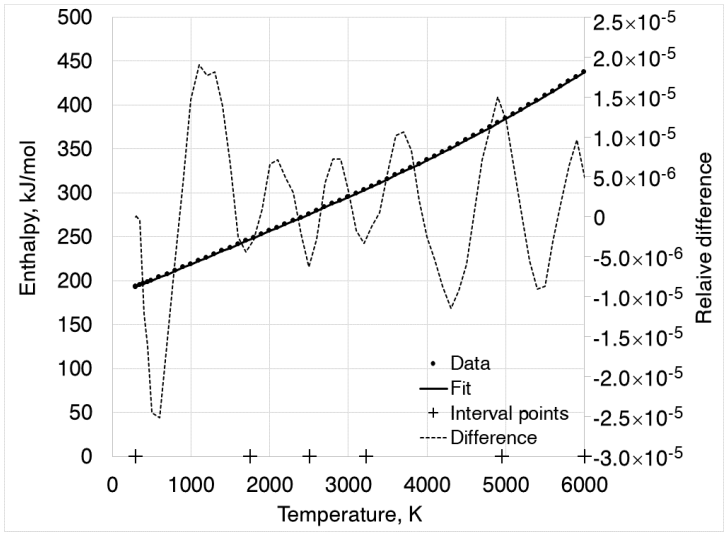


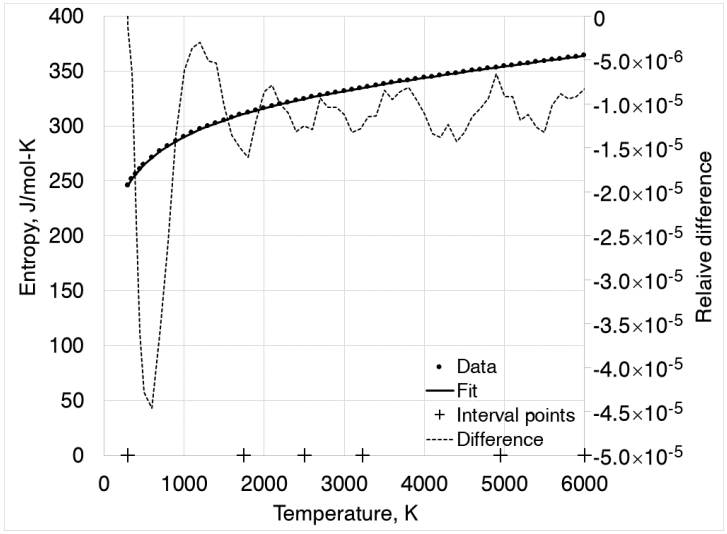


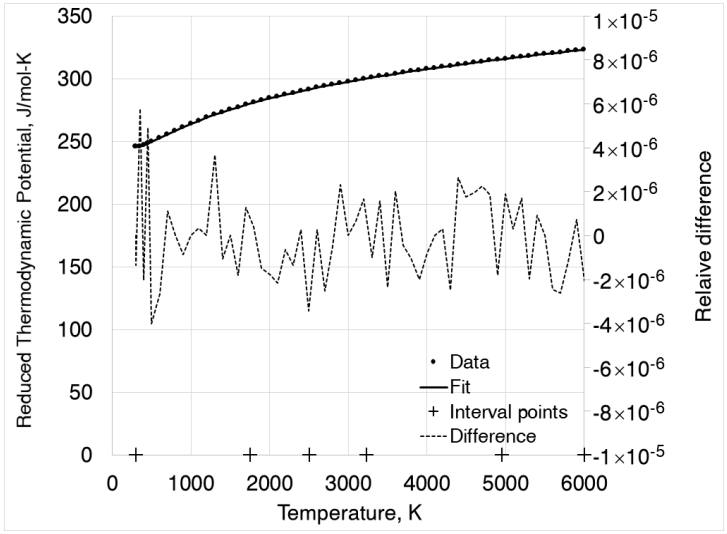
Cobalt Chloride (CoCl) – Cl1Co1(g)

MW = 94.384 g/mol – $\Delta_f H^o$(298.15 K) = 192.882 kJ/mol – S^o(298.15 K) = 245.679 J/mol-K – p^o = 0.1 MPa



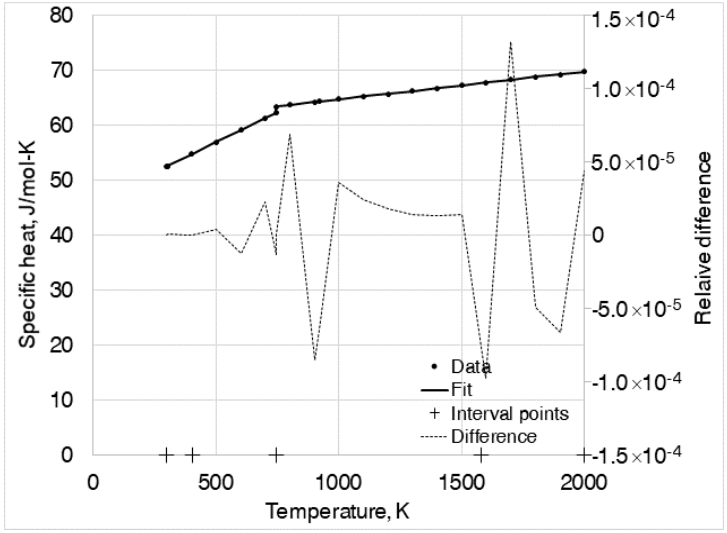


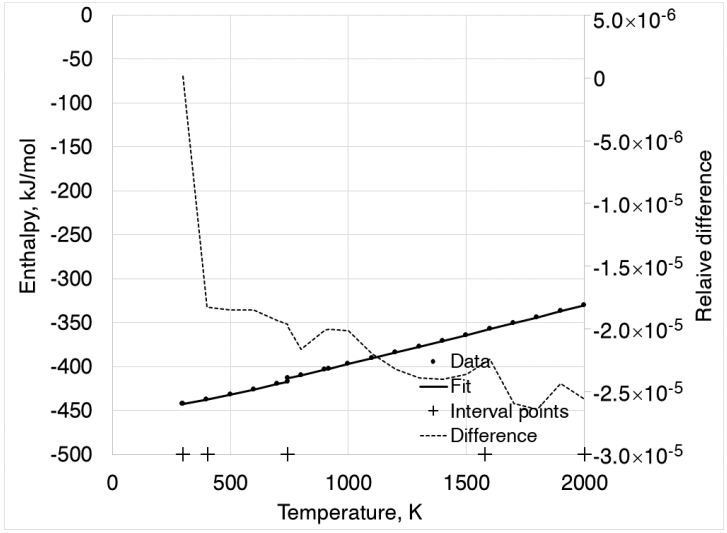


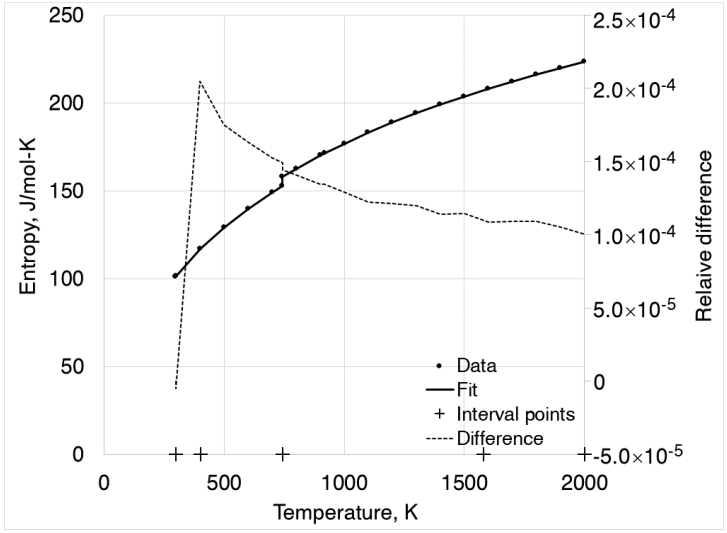


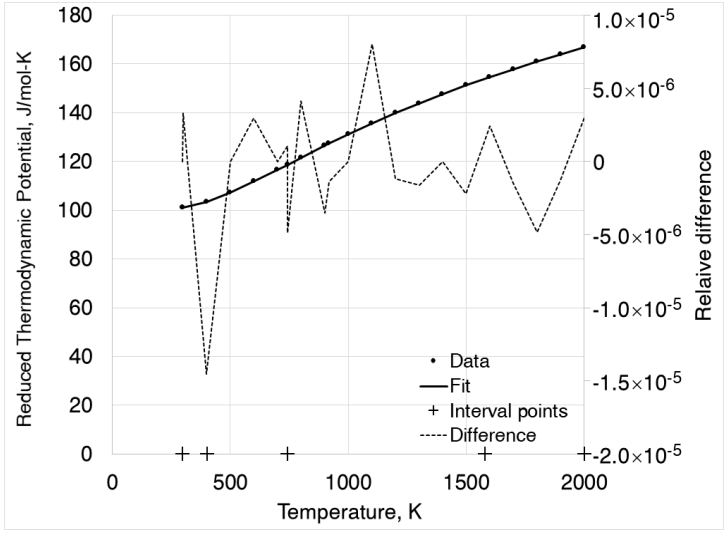
Cesium Chloride (CsCl) – Cl1Cs1(cr)

MW = 168.358 g/mol – $\Delta_f H^o$(298.15 K) = -442.835 kJ/mol – S^o(298.15 K) = 101.182 J/mol-K – p^o = 0.1 MPa



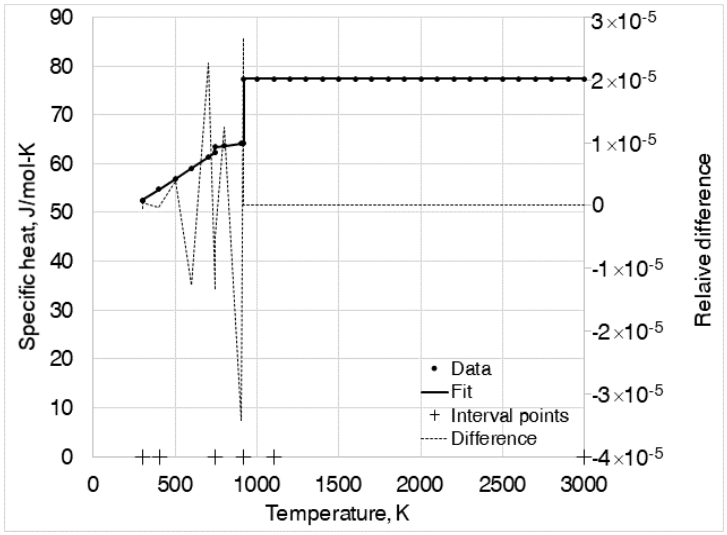


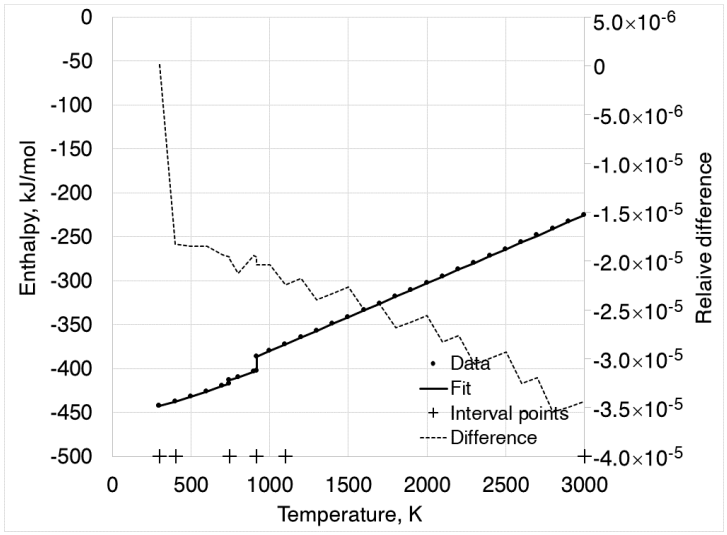


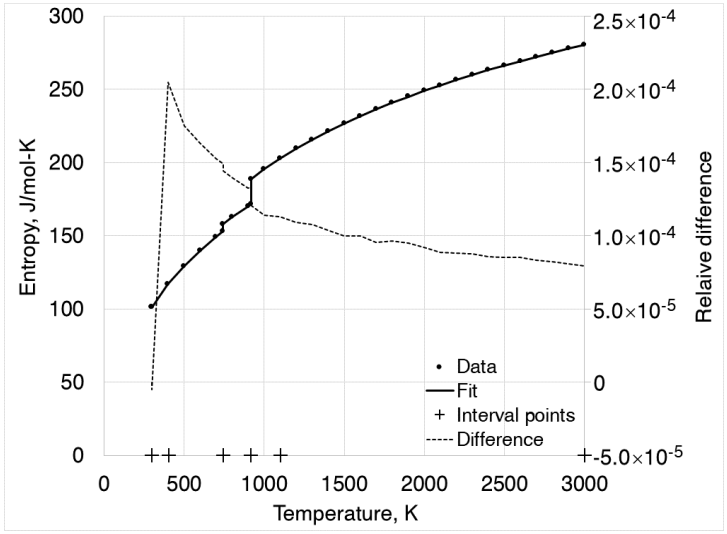


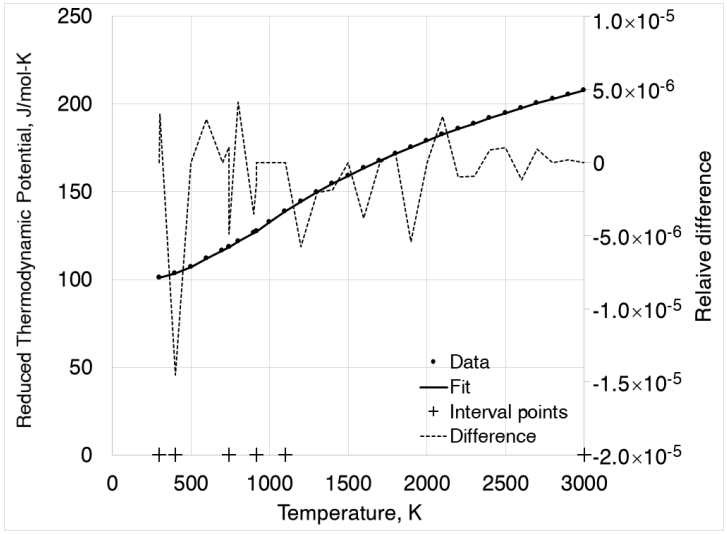
Cesium Chloride (CsCl) – Cl1Cs1(cr,l)

MW = 168.358 g/mol – $\Delta_f H^o$(298.15 K) = -442.835 kJ/mol – S^o(298.15 K) = 101.182 J/mol-K – p^o = 0.1 MPa



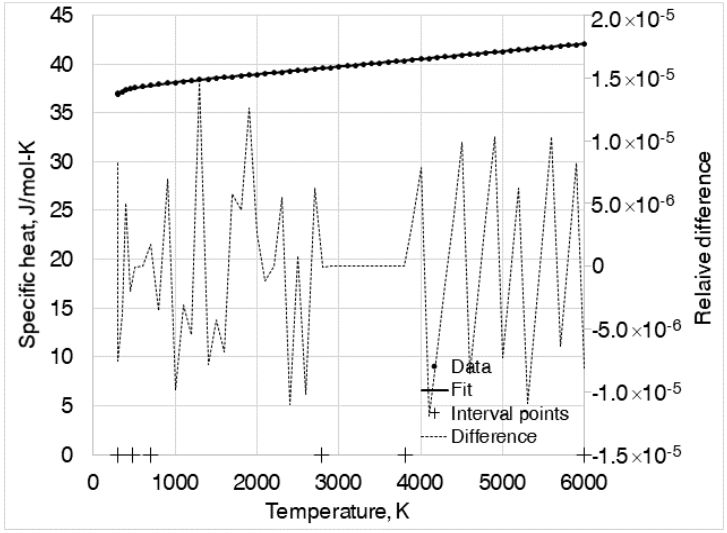


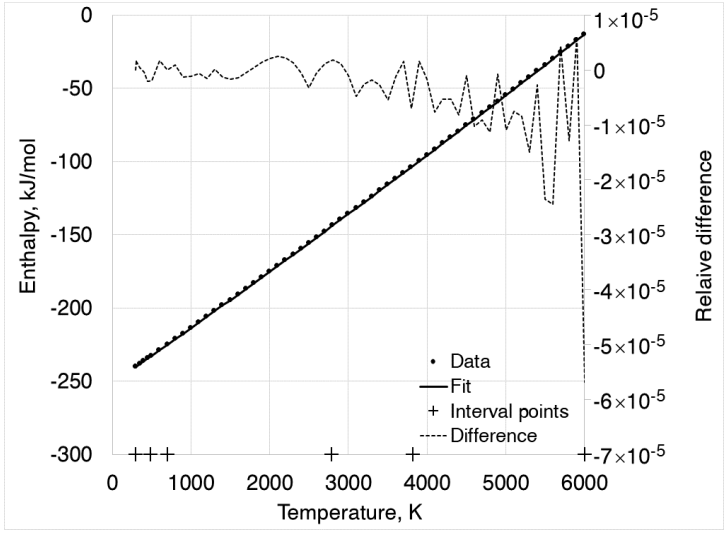


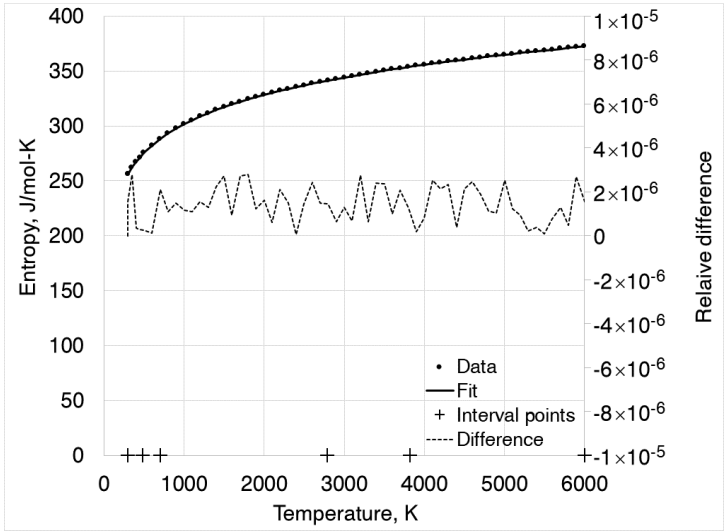


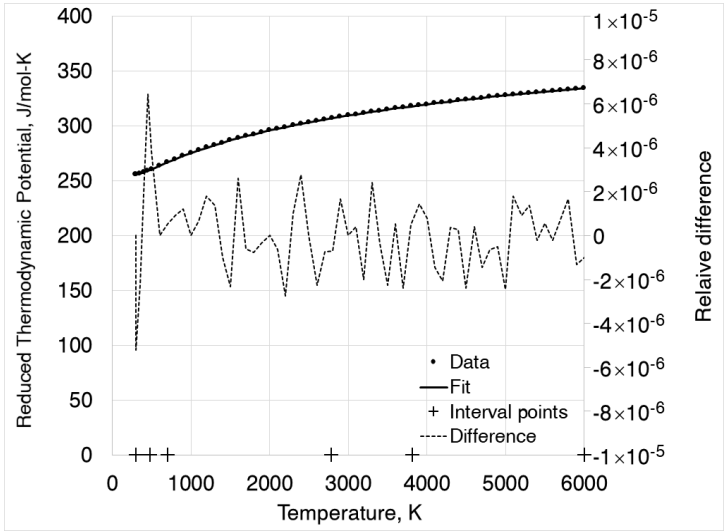
Cesium Chloride (CsCl) – Cl1Cs1(g)

MW = 168.358 g/mol – $\Delta_f H^o$(298.15 K) = -240.162 kJ/mol – S^o(298.15 K) = 256.07 J/mol-K – p^o = 0.1 MPa



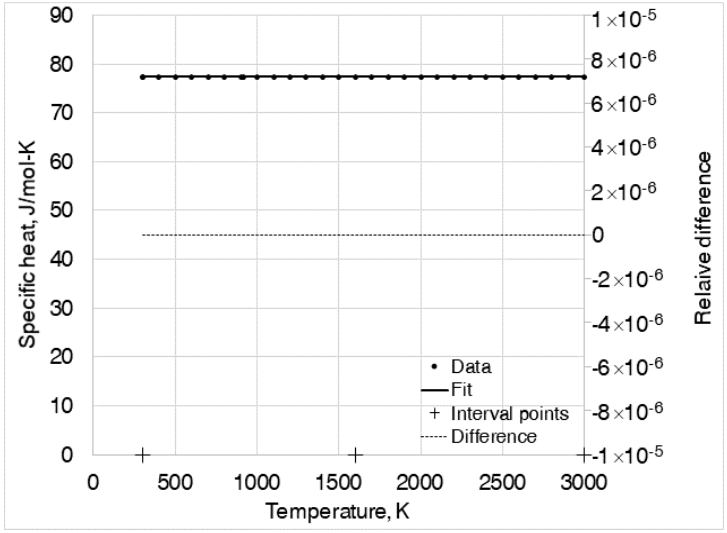


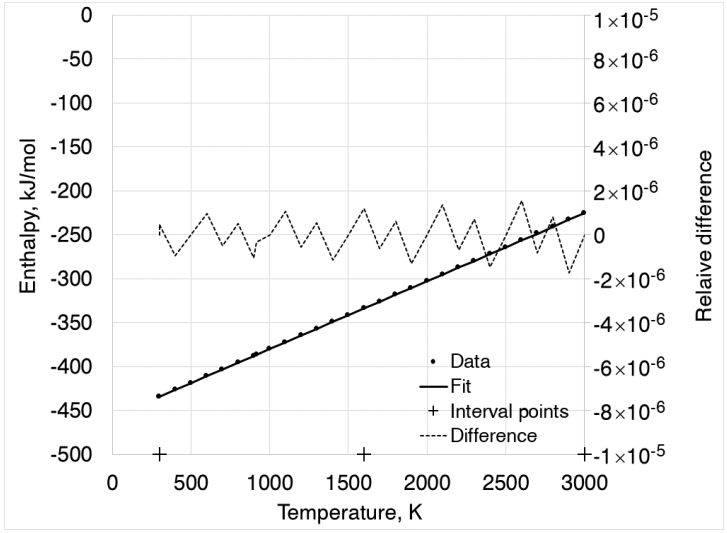


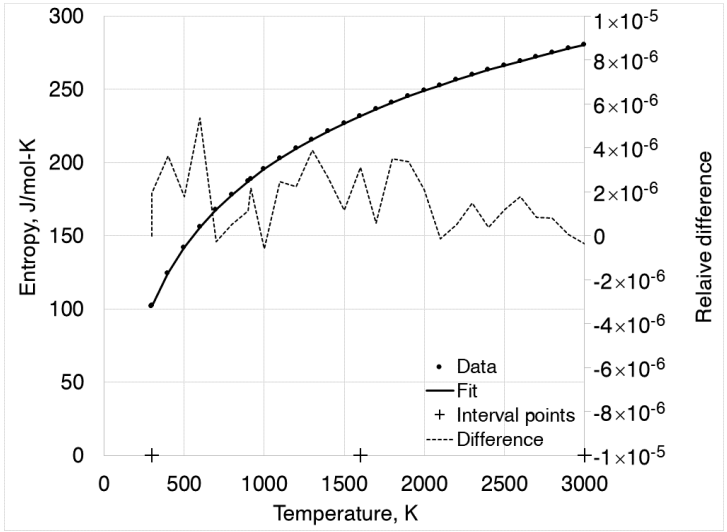


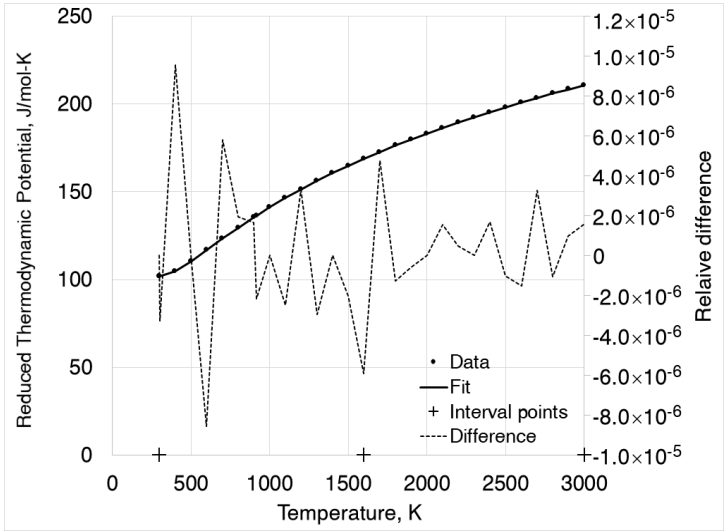
Cesium Chloride (CsCl) – Cl1Cs1(l)

MW = 168.358 g/mol – $\Delta_f H^o$(298.15 K) = -434.468 kJ/mol – S^o(298.15 K) = 101.708 J/mol-K – p^o = 0.1 MPa



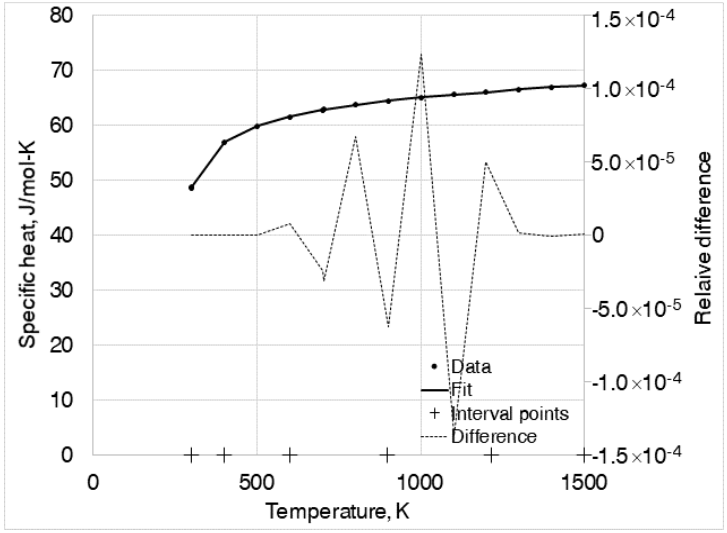


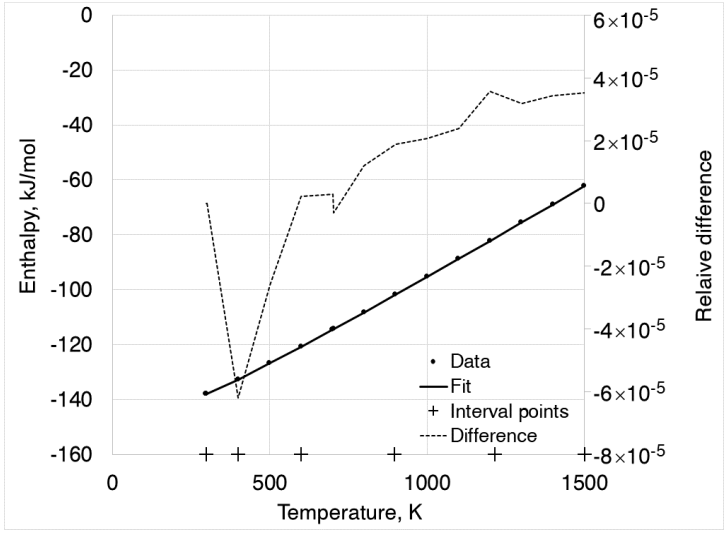


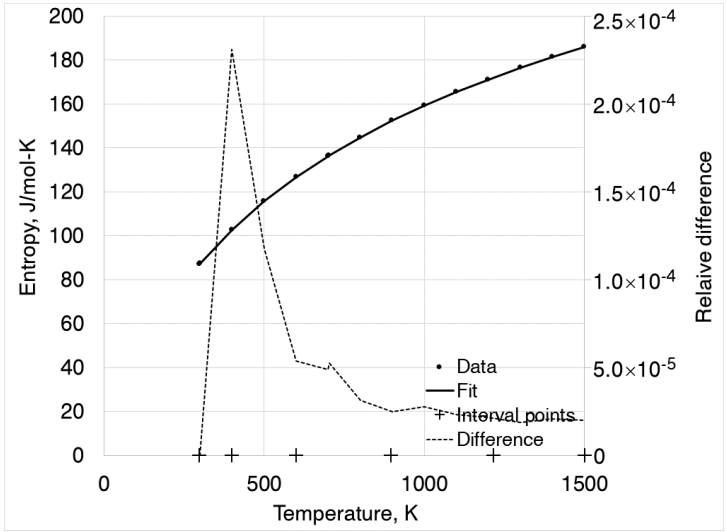


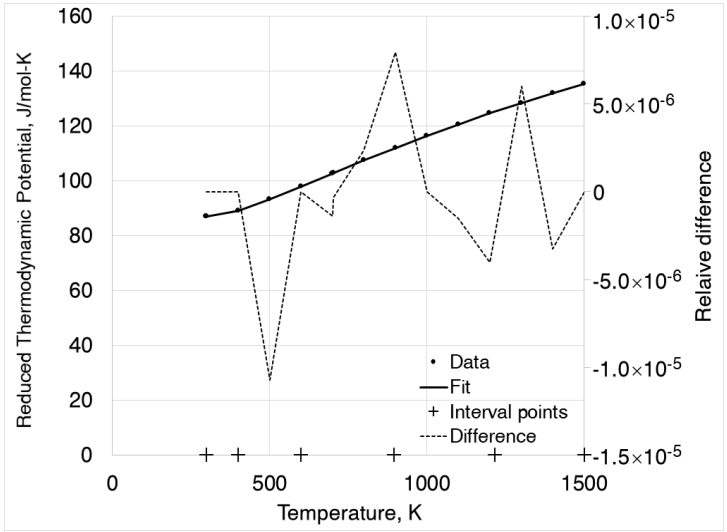
Copper Chloride (CuCl) – Cl1Cu1(cr)

MW = 98.999 g/mol – $\Delta_f H^o$(298.15 K) = -138.072 kJ/mol – S^o(298.15 K) = 87.027 J/mol-K – p^o = 0.1 MPa



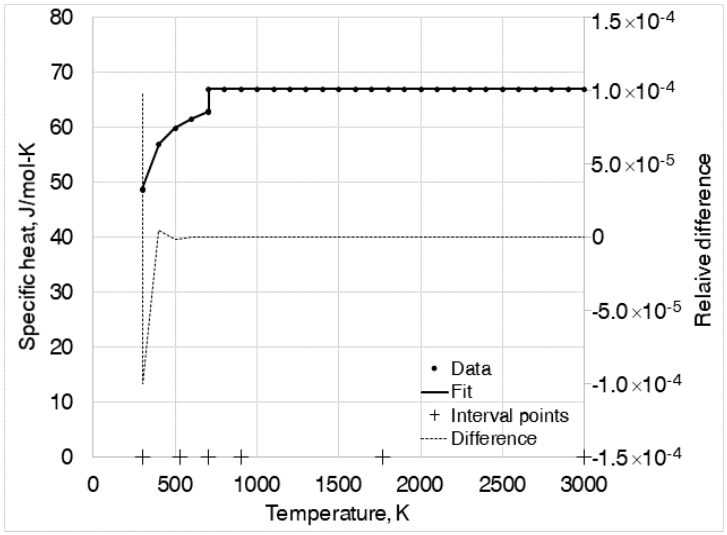


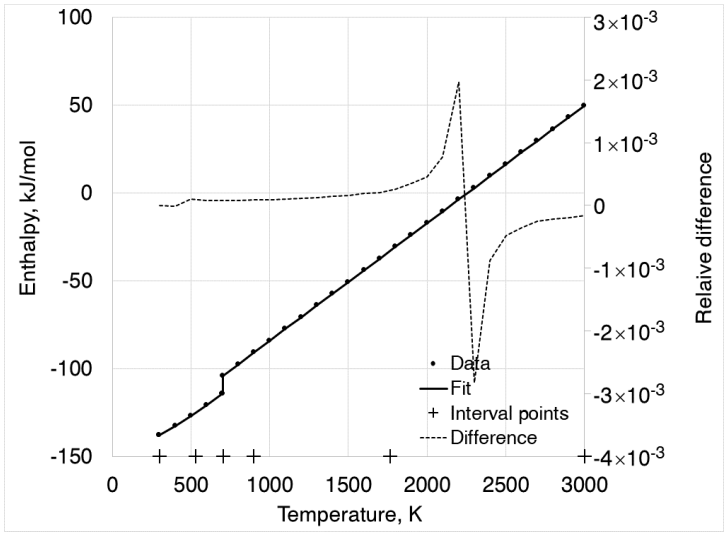


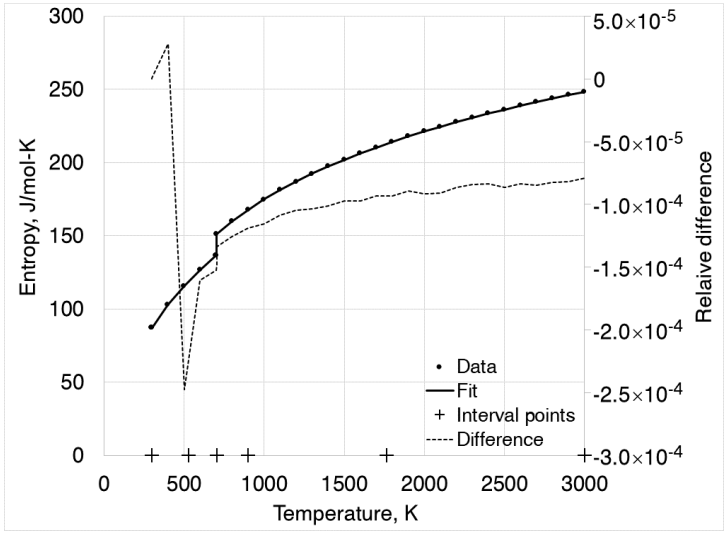


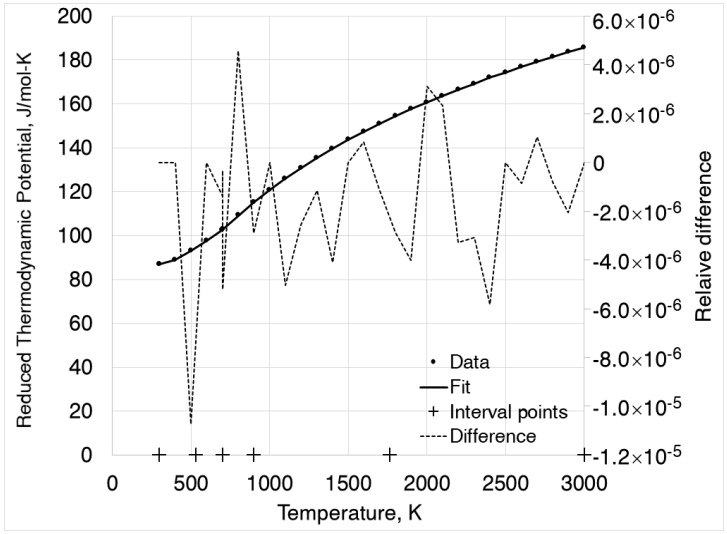
Copper Chloride (CuCl) – Cl1Cu1(cr,l)

MW = 98.999 g/mol – $\Delta_f H^o$(298.15 K) = -138.072 kJ/mol – S^o(298.15 K) = 87.027 J/mol-K – p^o = 0.1 MPa



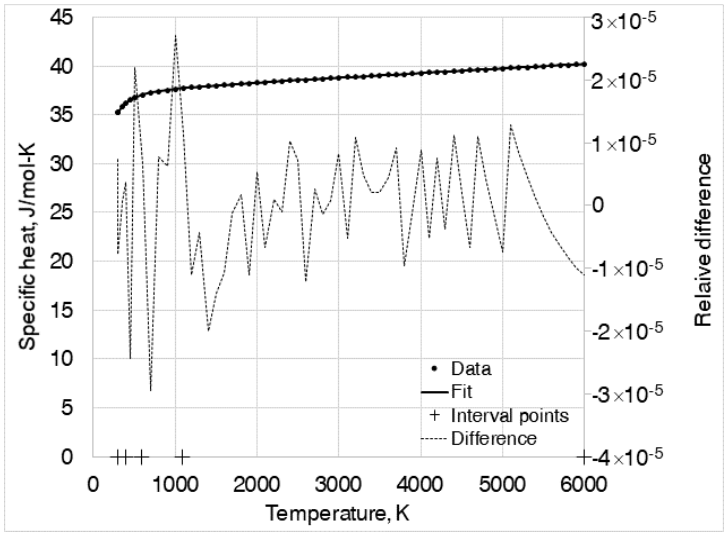


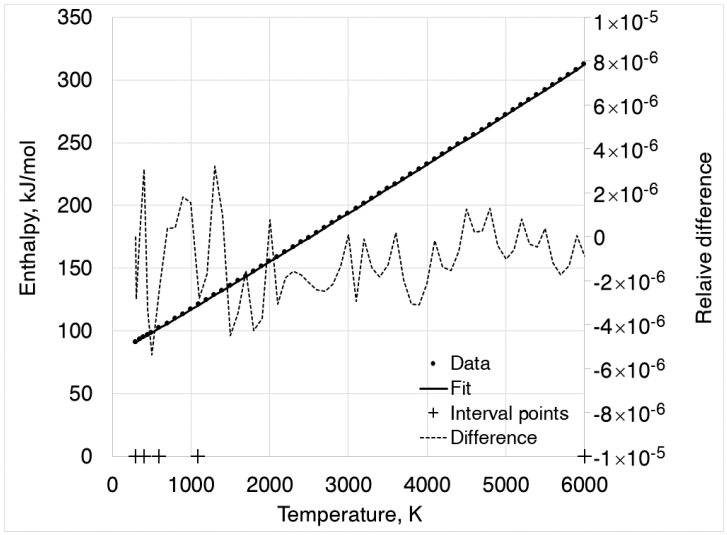


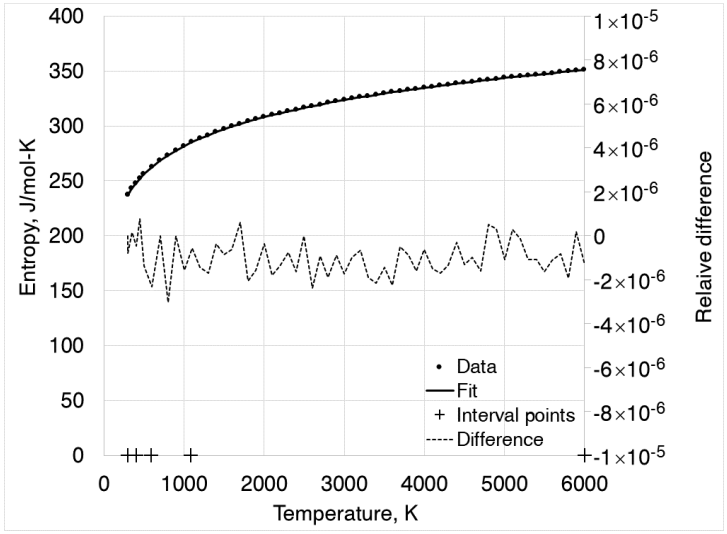


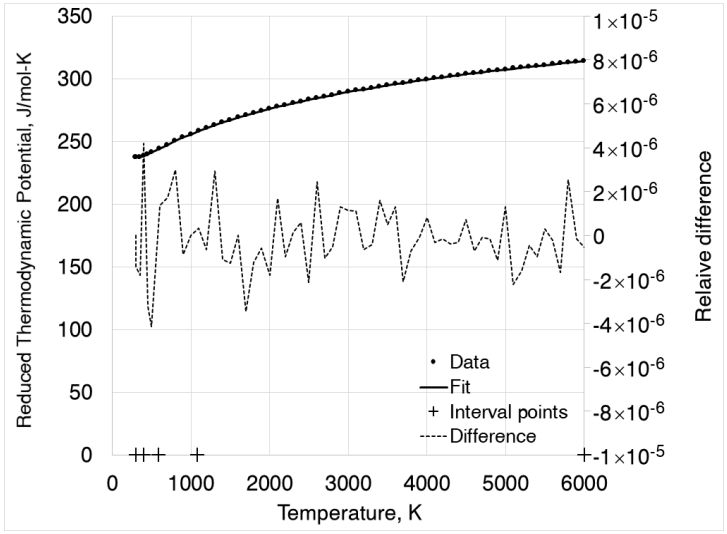
Copper Chloride (CuCl) – Cl1Cu1(g)

MW = 98.999 g/mol – $\Delta_f H^o$(298.15 K) = 91.086 kJ/mol – S^o(298.15 K) = 237.207 J/mol-K – p^o = 0.1 MPa



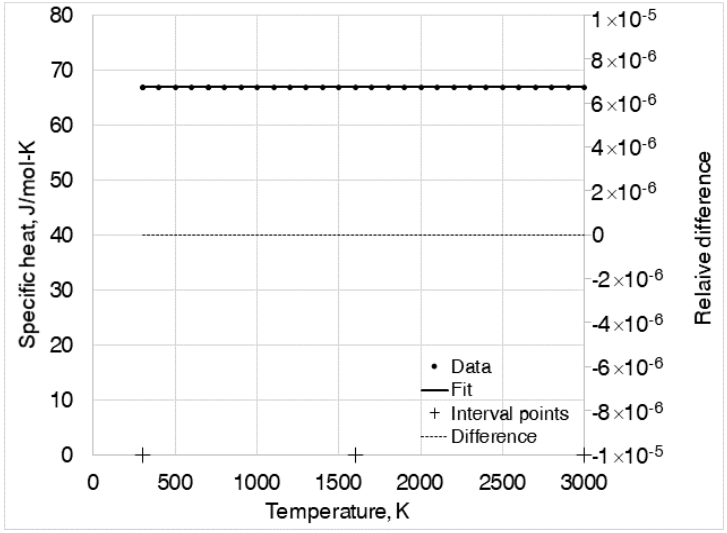


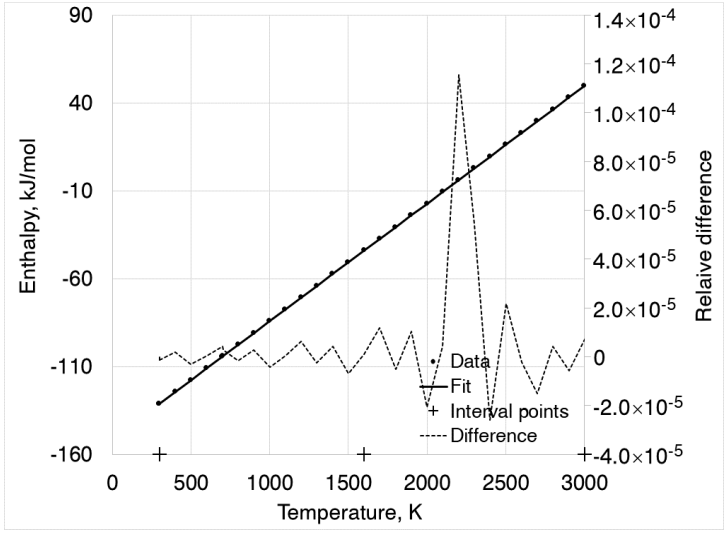


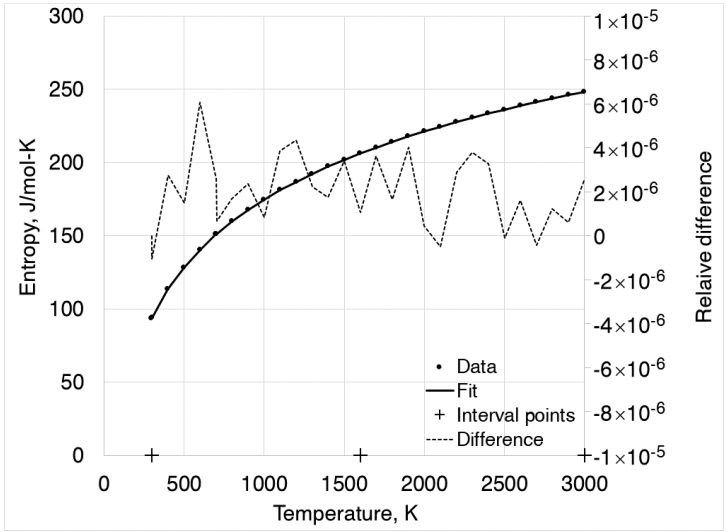


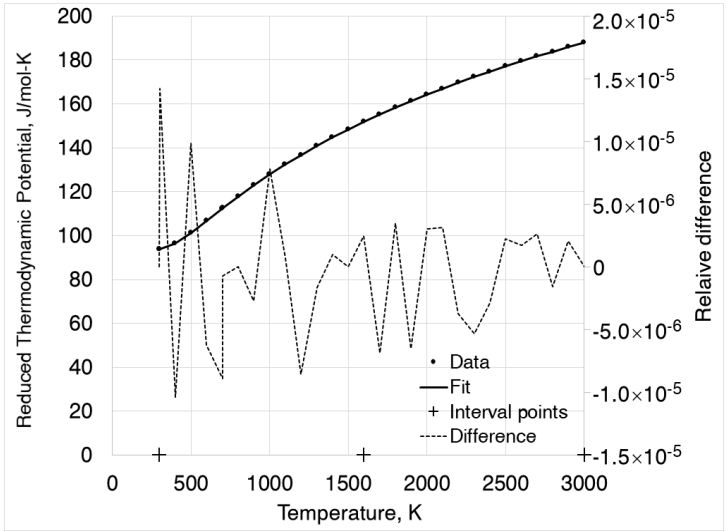
Copper Chloride (CuCl) – Cl1Cu1(l)

MW = 98.999 g/mol – $\Delta_f H^o$(298.15 K) = -131.178 kJ/mol – S^o(298.15 K) = 93.752 J/mol-K – p^o = 0.1 MPa



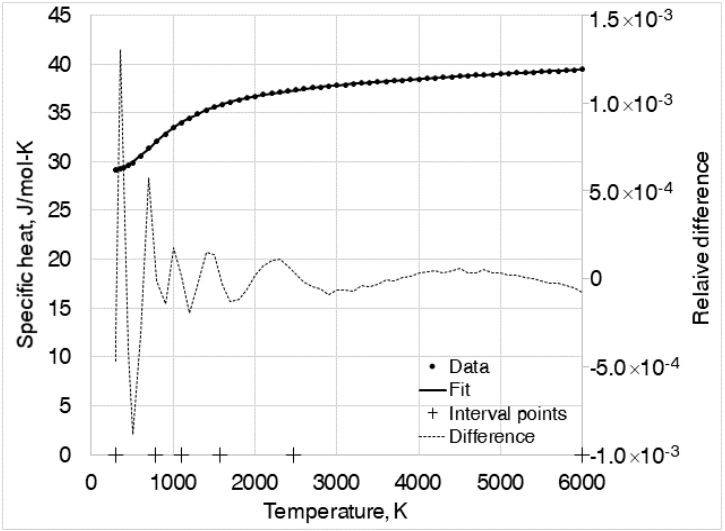


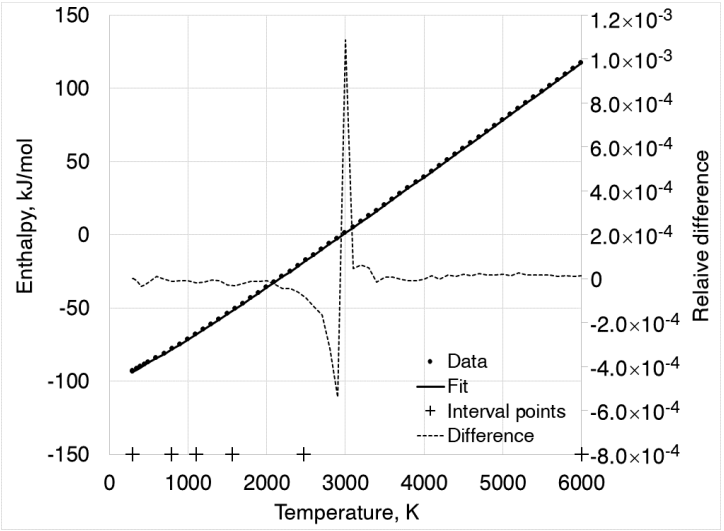


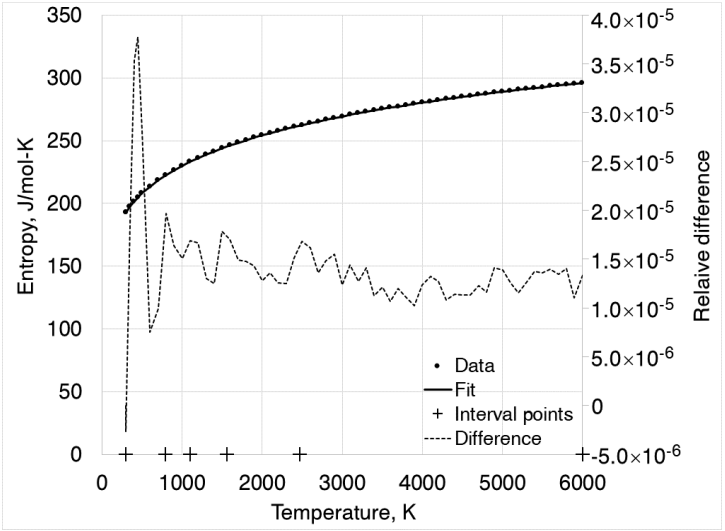


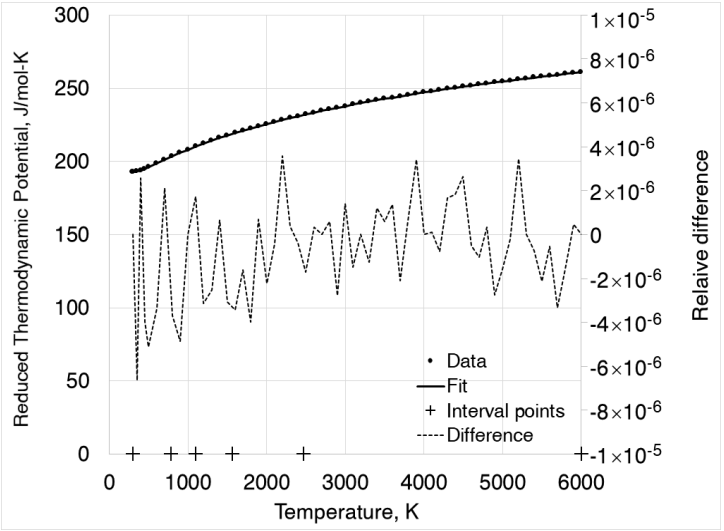
Hydrochloric Acid-D (DCl) – Cl1D1(g)

MW = 329 g/mol – $\Delta_f H^o$(298.15 K) = -93.345 kJ/mol – S^o(298.15 K) = 192.741 J/mol-K – p^o = 0.1 MPa



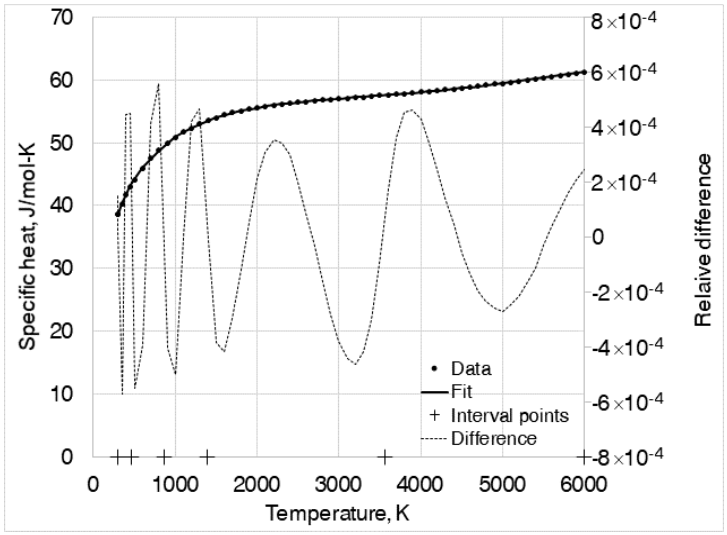


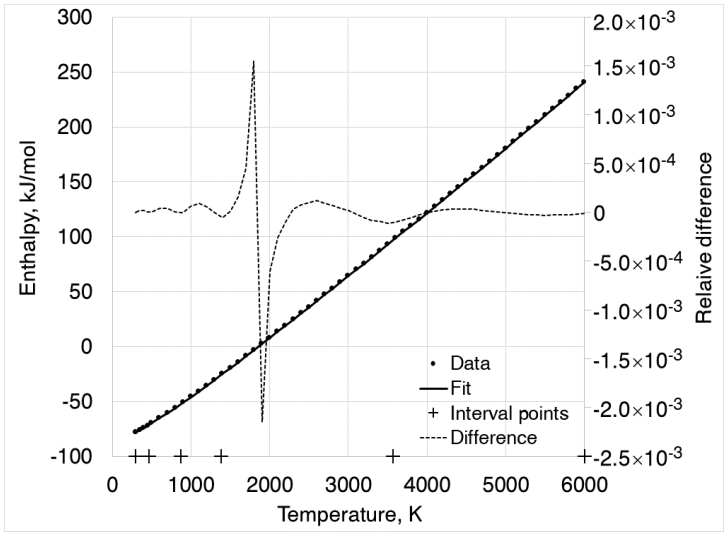


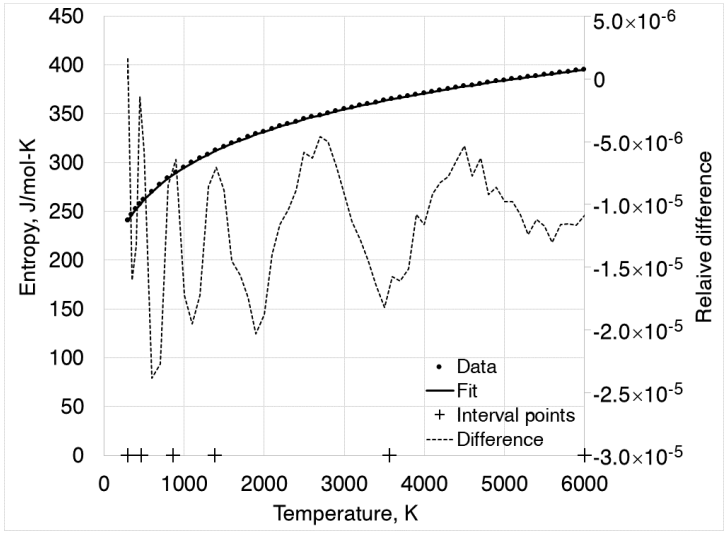


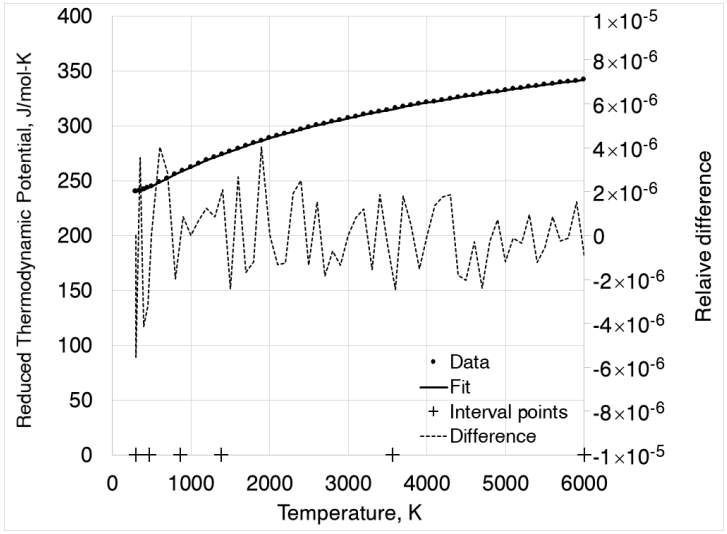
Hypochlorous Acid-D (DOCl) – Cl1D1O1(g)

MW = 345 g/mol – $\Delta_f H^o$(298.15 K) = -78.241 kJ/mol – S^o(298.15 K) = 240.246 J/mol-K – p^o = 0.1 MPa



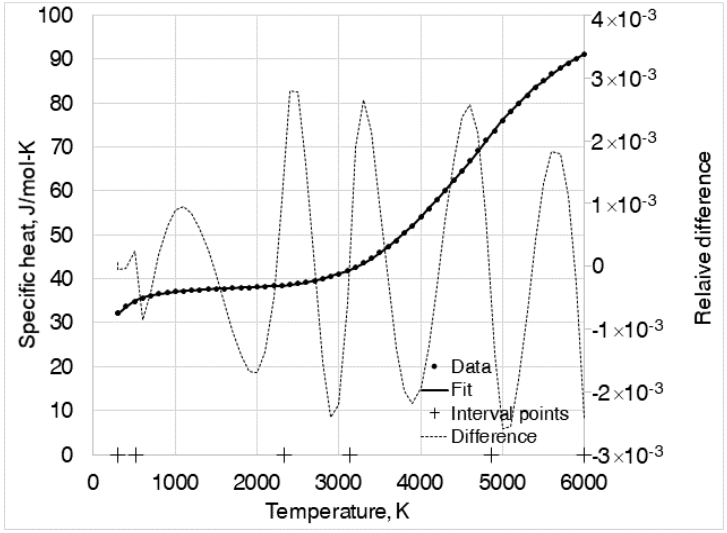


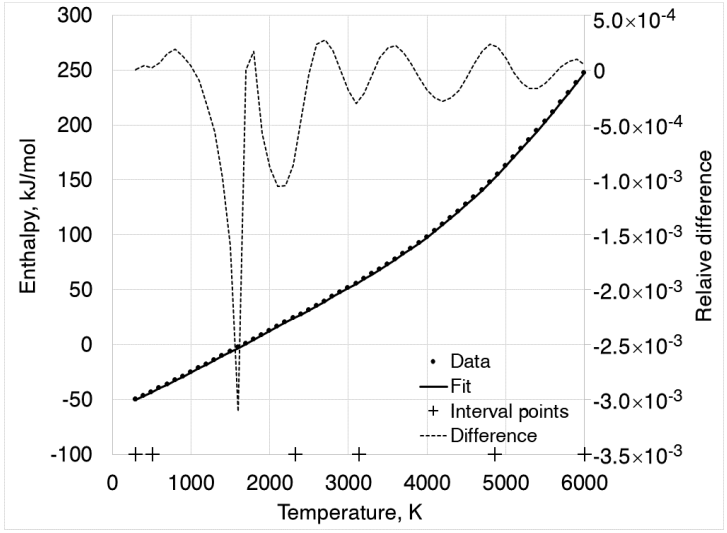


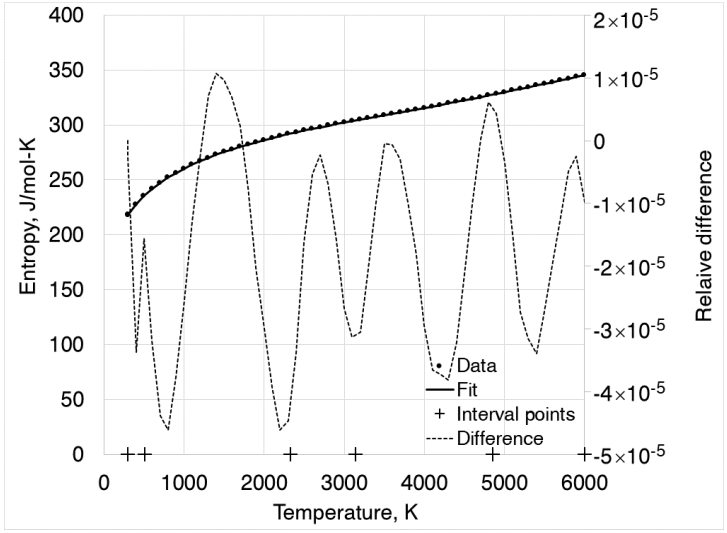


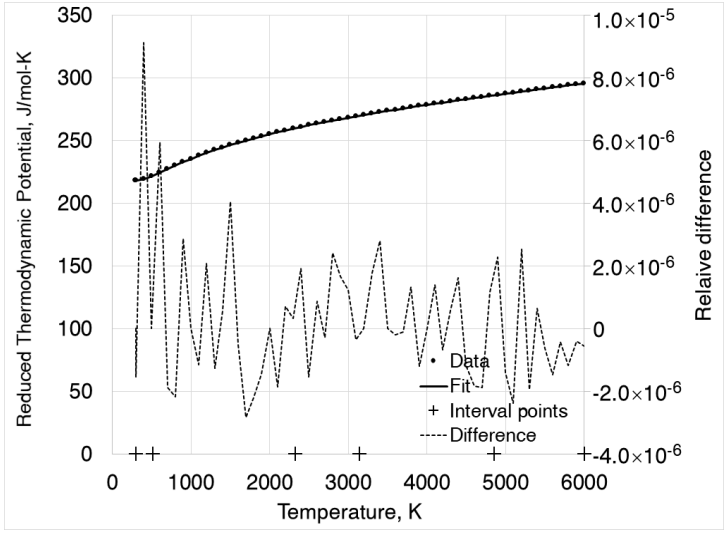
Chlorine Fluoride (ClF) – Cl1F1(g)

MW = 54.451 g/mol – $\Delta_f H^o$(298.15 K) = -50.292 kJ/mol – S^o(298.15 K) = 217.938 J/mol-K – p^o = 0.1 MPa



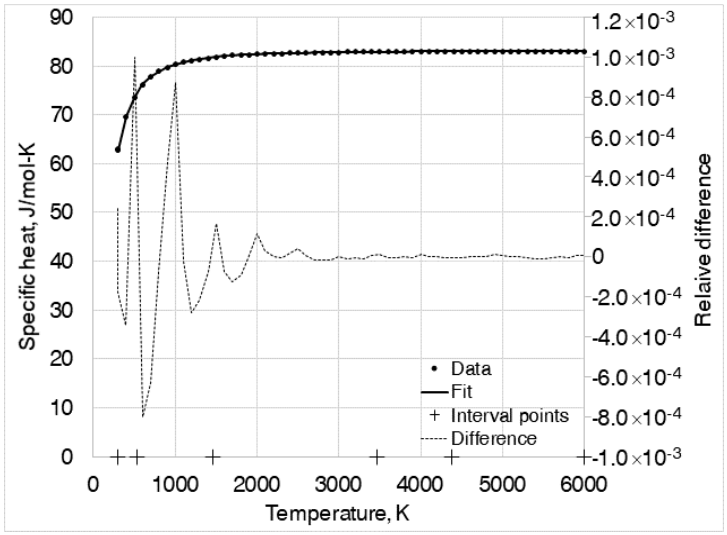


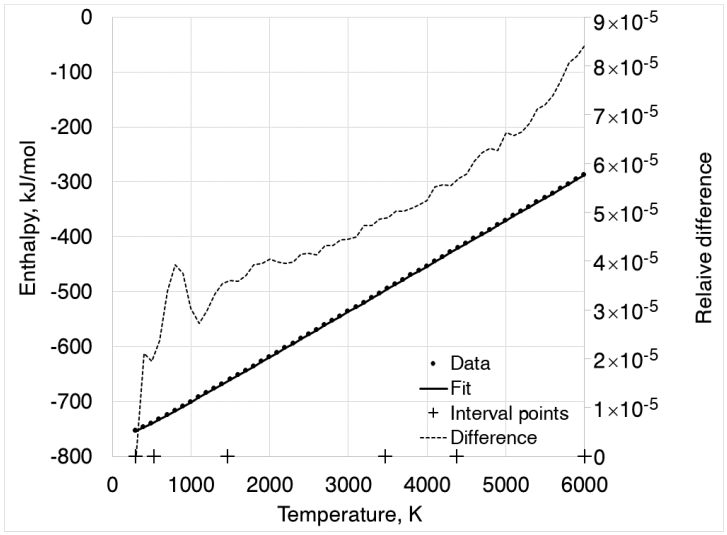


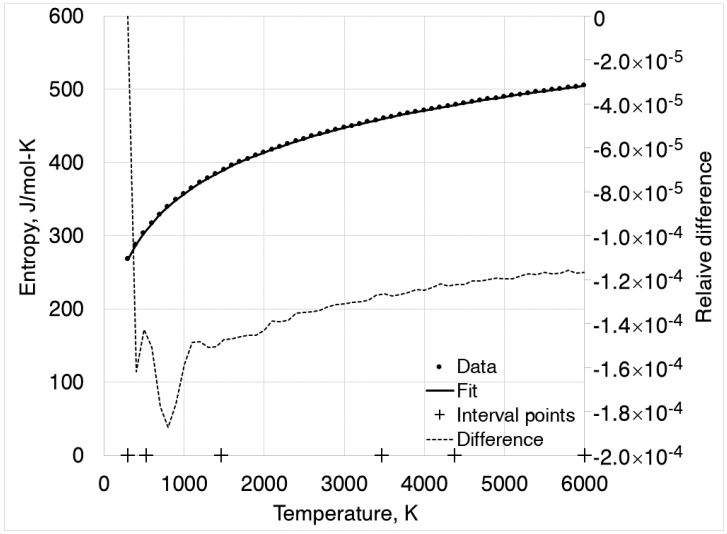


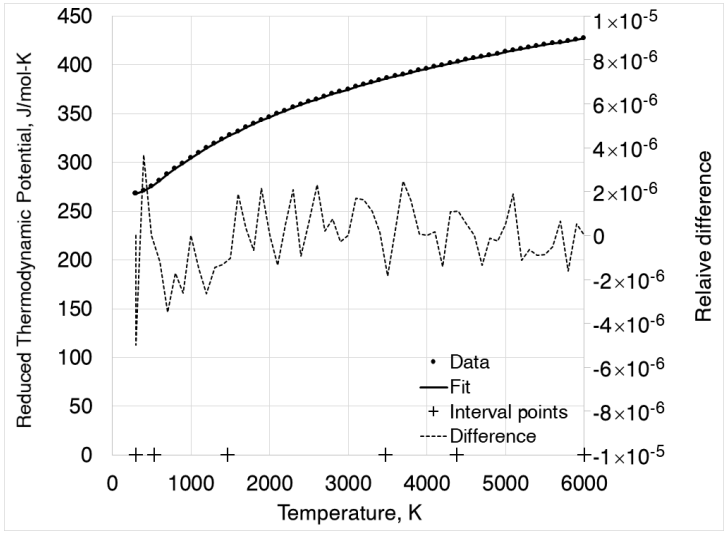
Lithium Chloride Fluoride (Li_2ClF) – Cl1F1Li2(g)

MW = 68.333 g/mol – $\Delta_f H^o$(298.15 K) = -753.957 kJ/mol – S^o(298.15 K) = 268.04 J/mol-K – p^o = 0.1 MPa



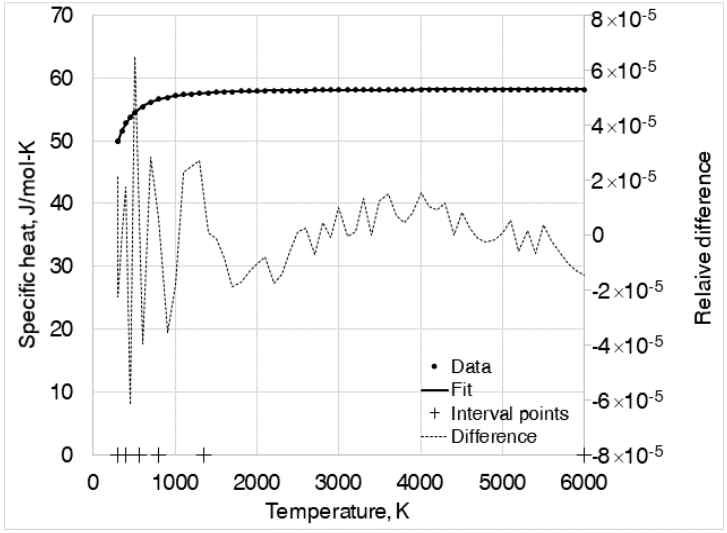


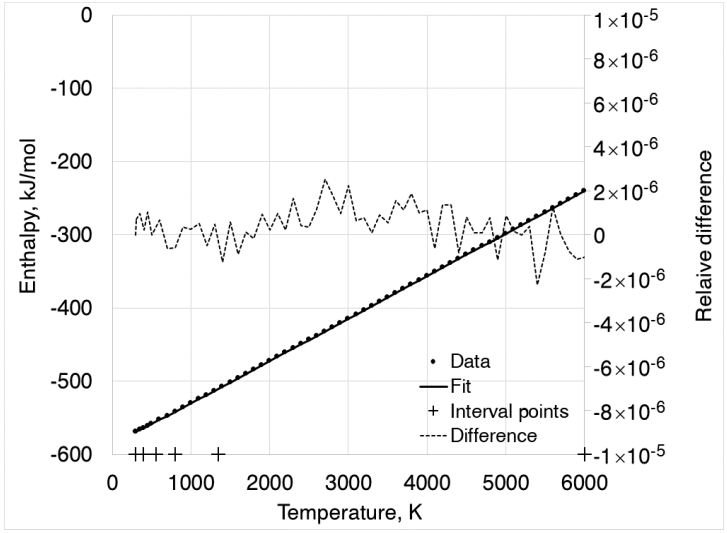


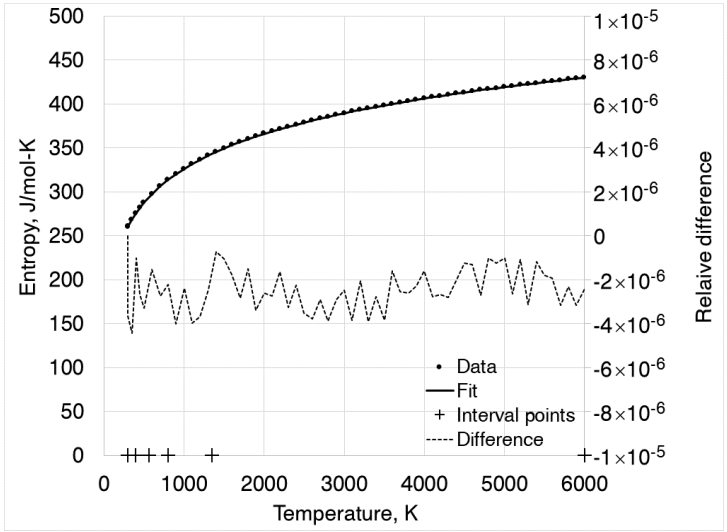


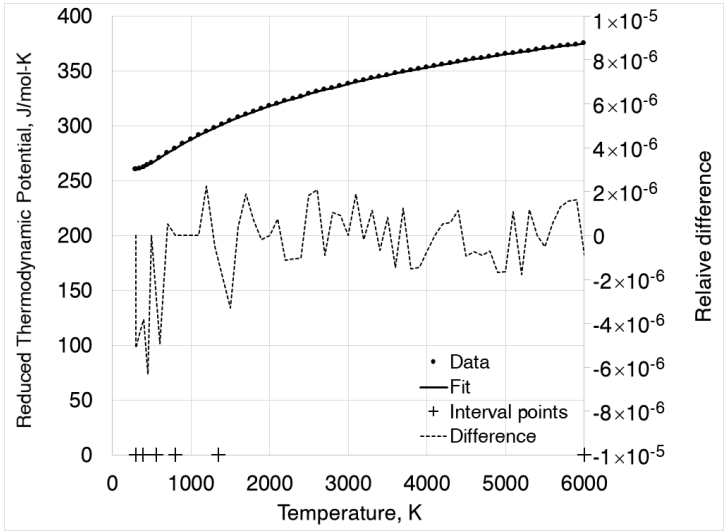
Magnesium Chloride Fluoride (MgClF) – Cl1F1Mg1(g)

MW = 78.756 g/mol – $\Delta_f H^o$(298.15 K) = -569.024 kJ/mol – S^o(298.15 K) = 260.227 J/mol-K – p^o = 0.1 MPa



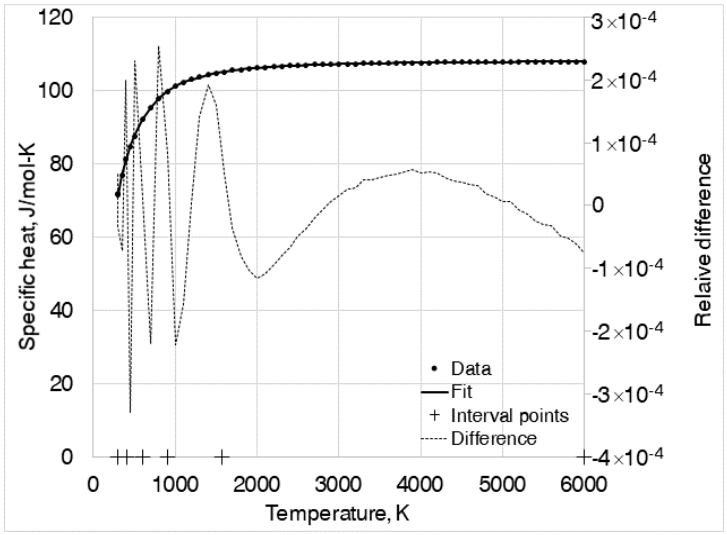


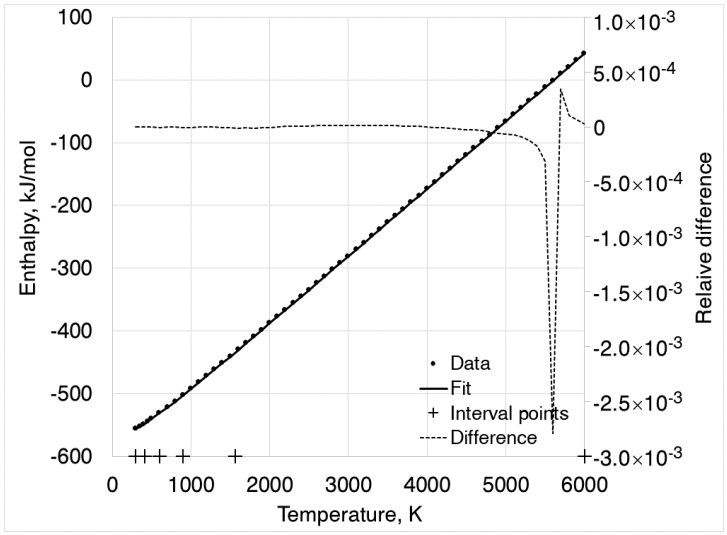


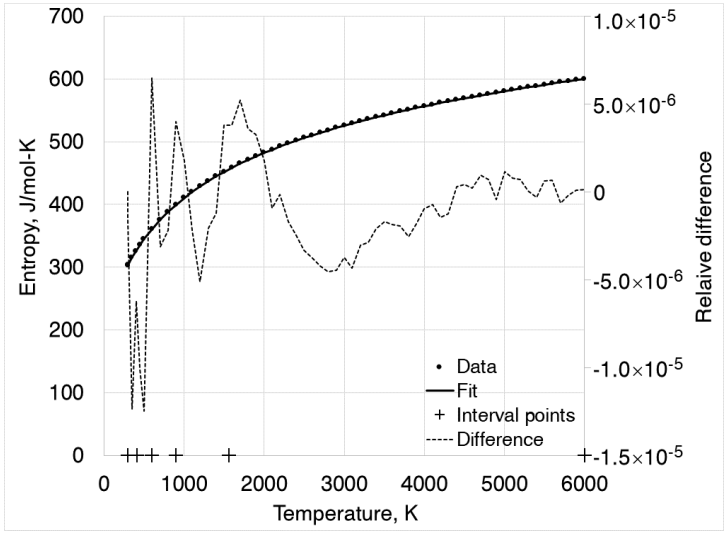


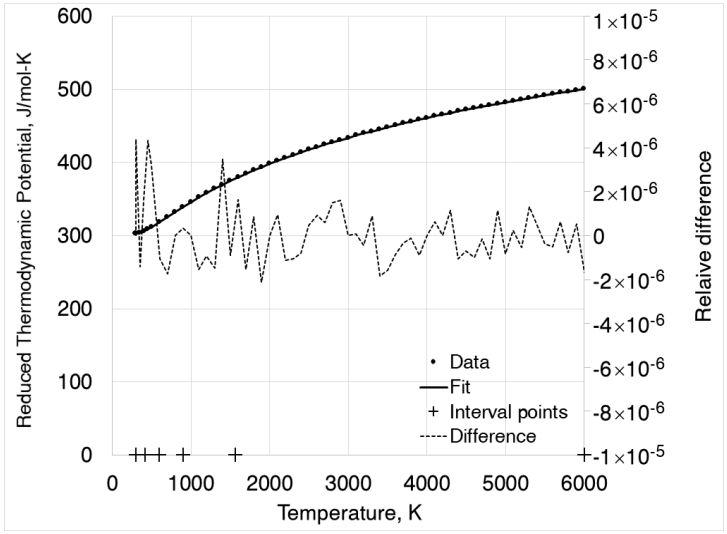
Sulfuryl Chloride Fluoride (SO_2ClF) – Cl1F1O2S1(g)

MW = 118.51 g/mol – $\Delta_f H^o$(298.15 K) = -556.472 kJ/mol – S^o(298.15 K) = 302.849 J/mol-K – p^o = 0.1 MPa



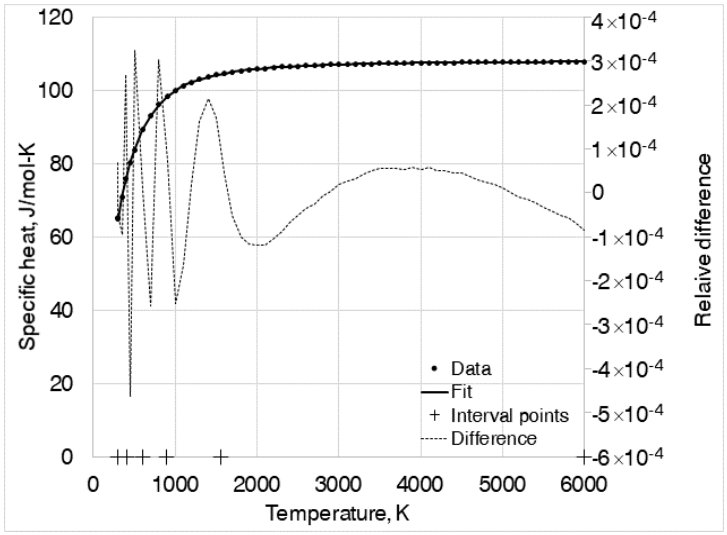


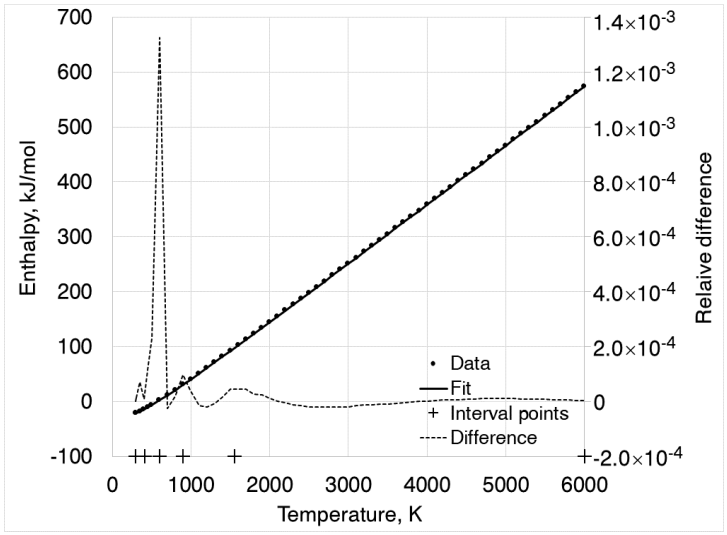


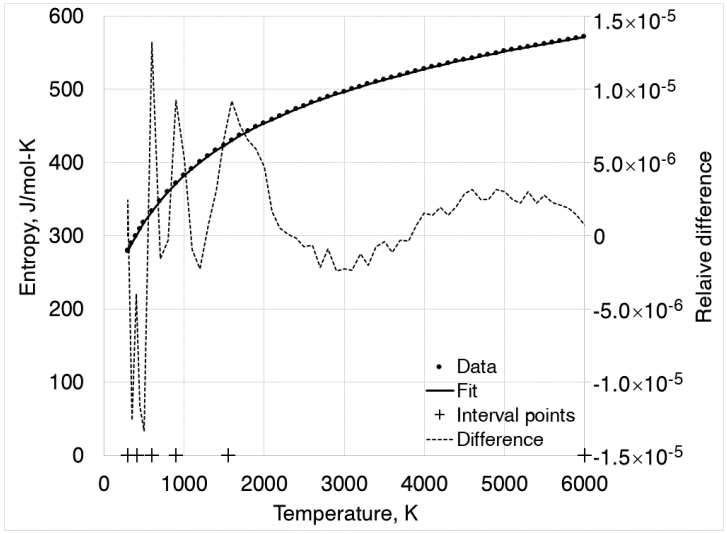


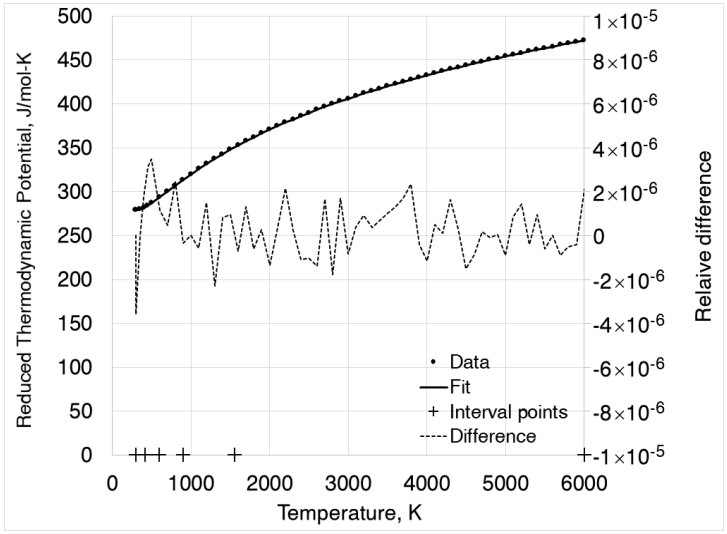
Perchloryl Fluoride (ClO_3F) – Cl1F1O3(g)

MW = 102.45 g/mol – $\Delta_f H^o$(298.15 K) = -21.422 kJ/mol – S^o(298.15 K) = 278.984 J/mol-K – p^o = 0.1 MPa



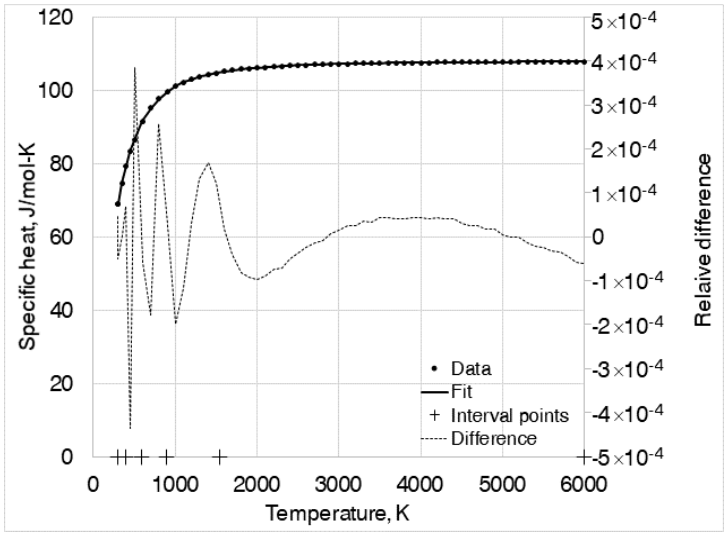


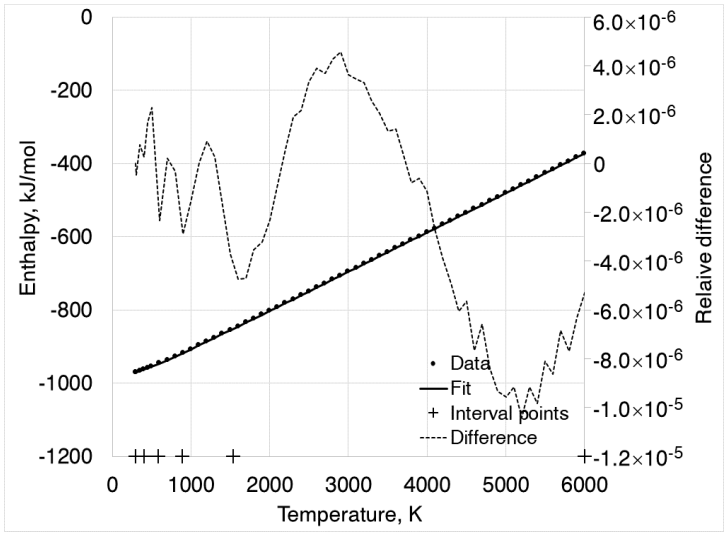


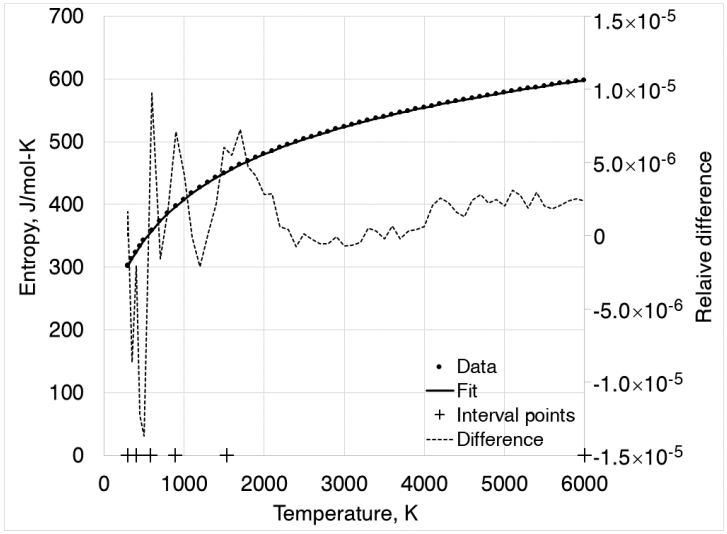


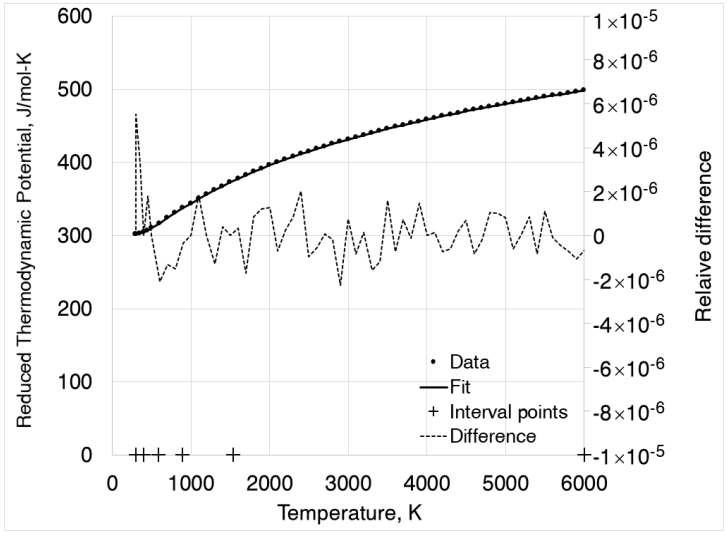
Phosphoryl Chloride Fluoride ($POClF_2$) – Cl1F2O1P1(g)

MW = 120.423 g/mol – $\Delta_f H^o$(298.15 K) = -970.688 kJ/mol – S^o(298.15 K) = 301.689 J/mol-K – p^o = 0.1 MPa



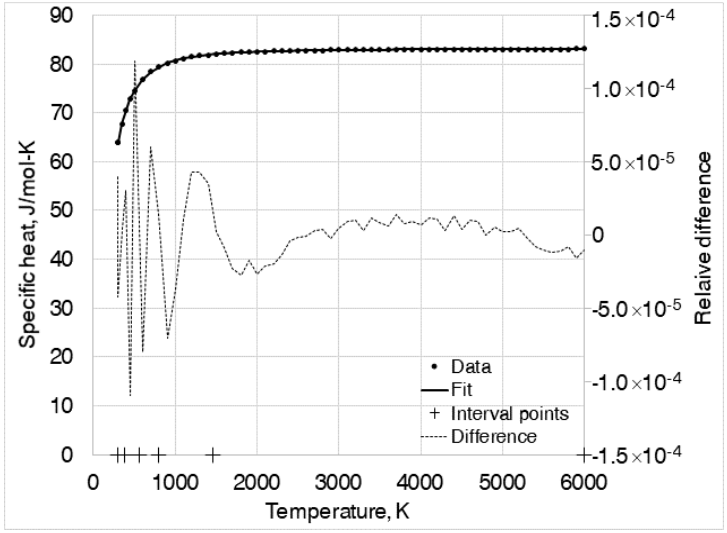


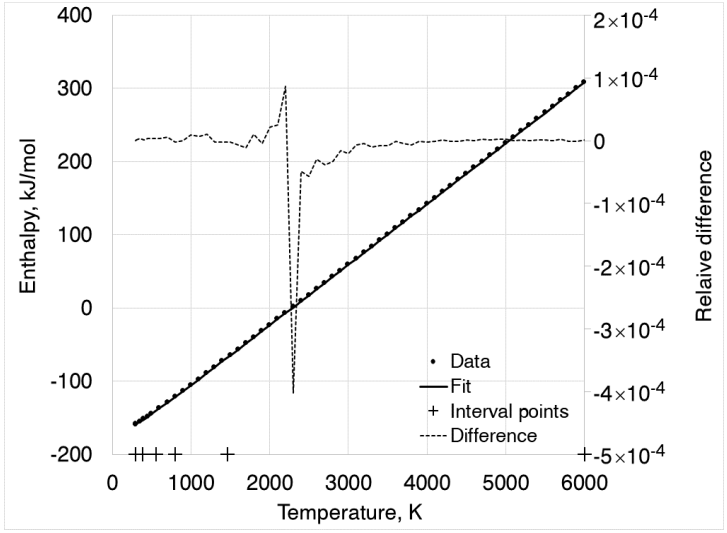


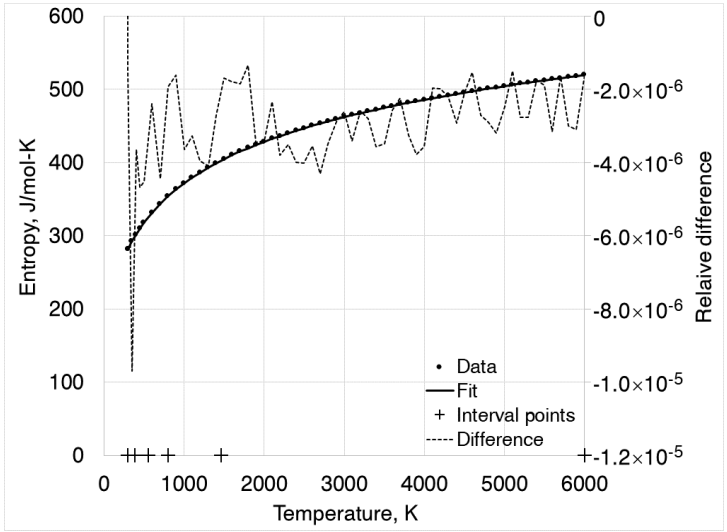


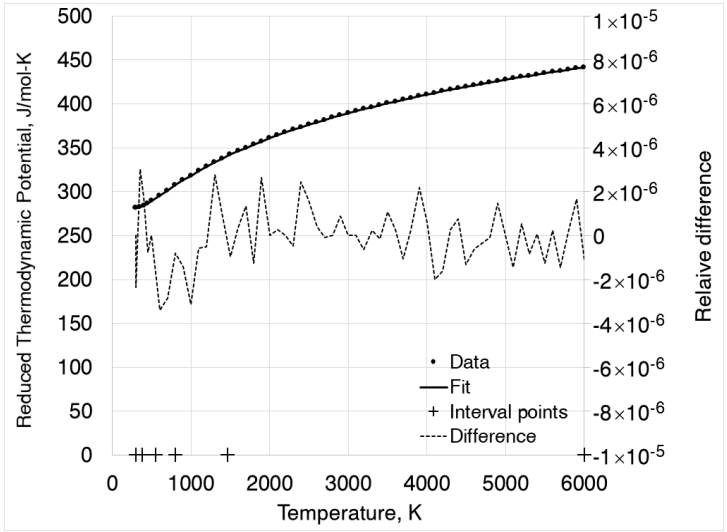
Chlorine Fluoride (ClF_3) – Cl1F3(g)

MW = 92.448 g/mol – $\Delta_f H^o$(298.15 K) = -158.866 kJ/mol – S^o(298.15 K) = 281.6 J/mol-K – p^o = 0.1 MPa



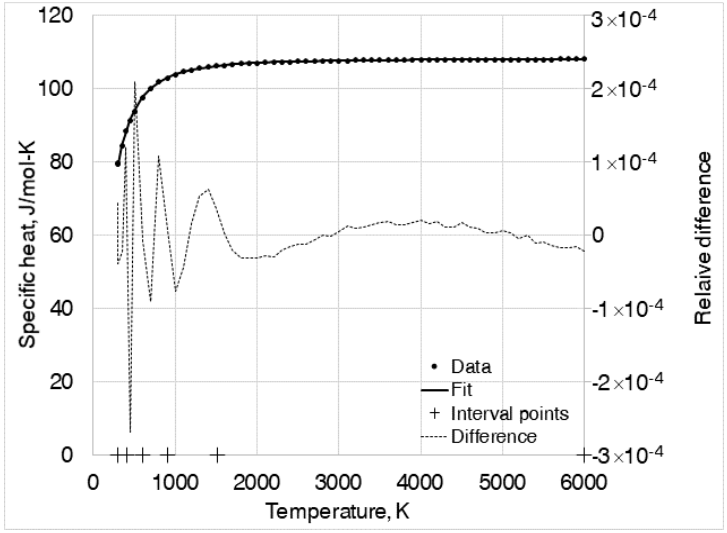


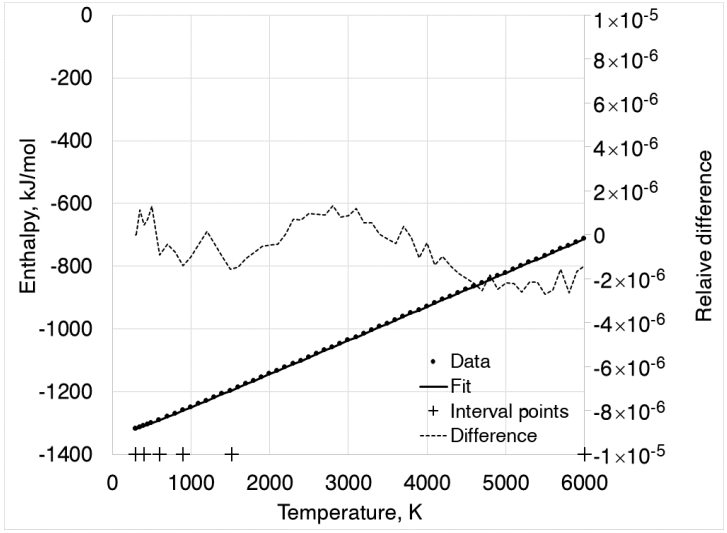


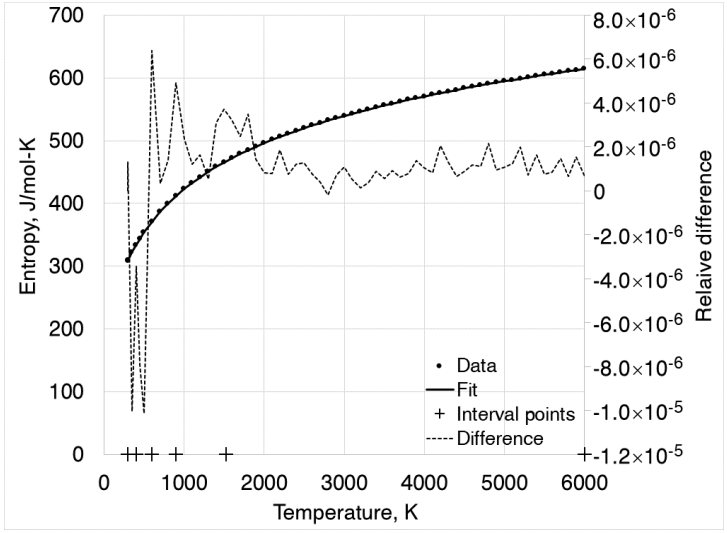


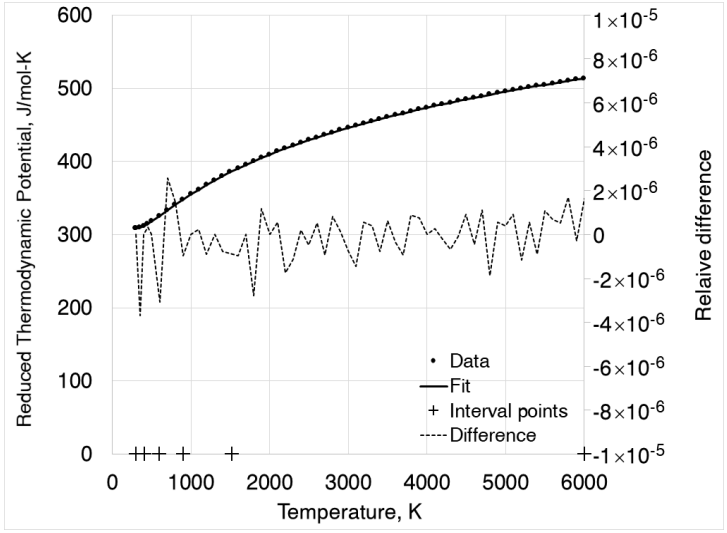
Chlorotrifluorosilane ($SiClF_3$) – Cl1F3Si1(g)

MW = 120.534 g/mol – $\Delta_f H^o$(298.15 K) = -1317.96 kJ/mol – S^o(298.15 K) = 308.679 J/mol-K – p^o = 0.1 MPa



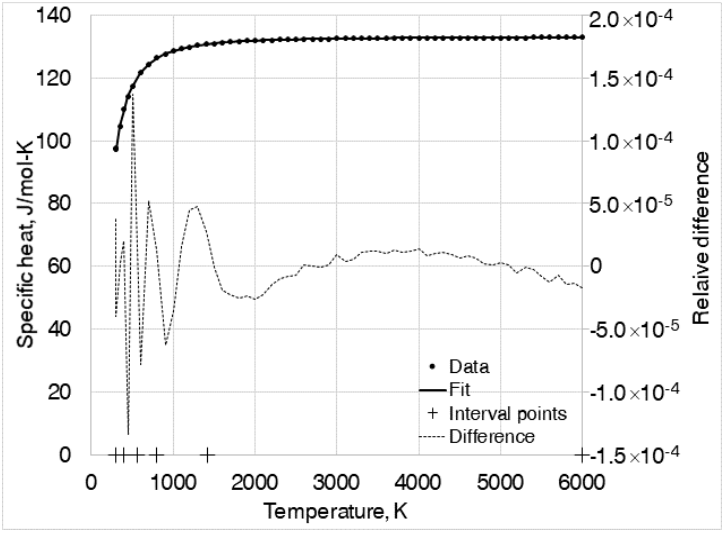


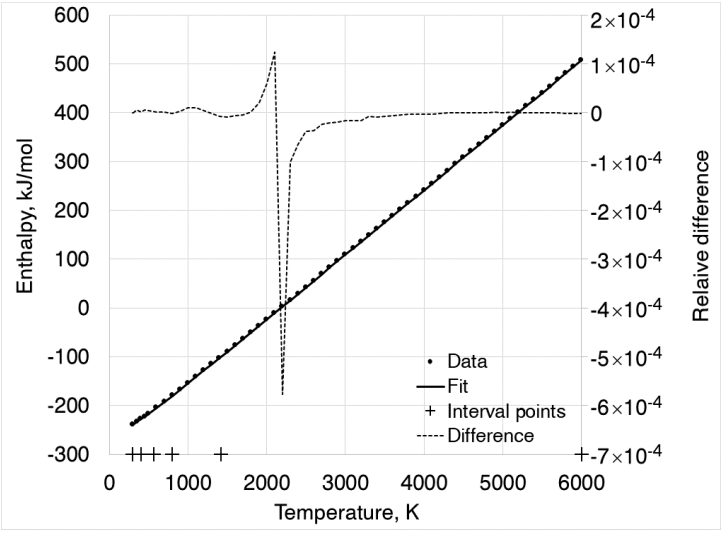


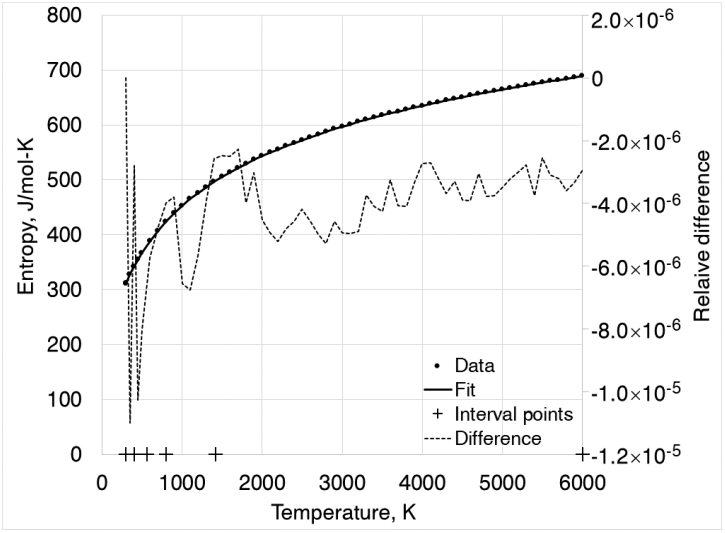


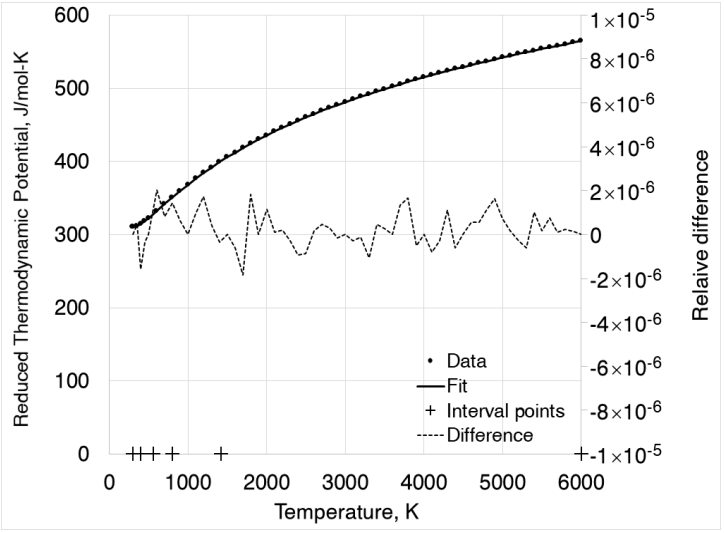
Chlorine Fluoride (ClF_5) – Cl1F5(g)

MW = 130.445 g/mol – $\Delta_f H^o$(298.15 K) = -238.488 kJ/mol – S^o(298.15 K) = 310.739 J/mol-K – p^o = 0.1 MPa



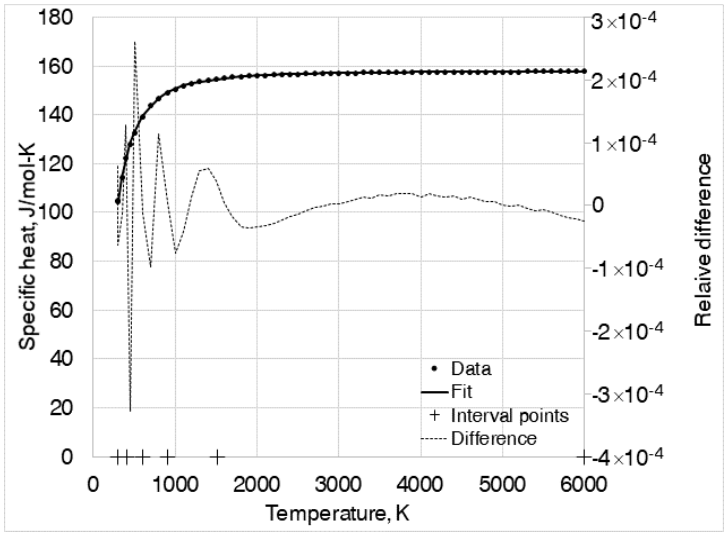


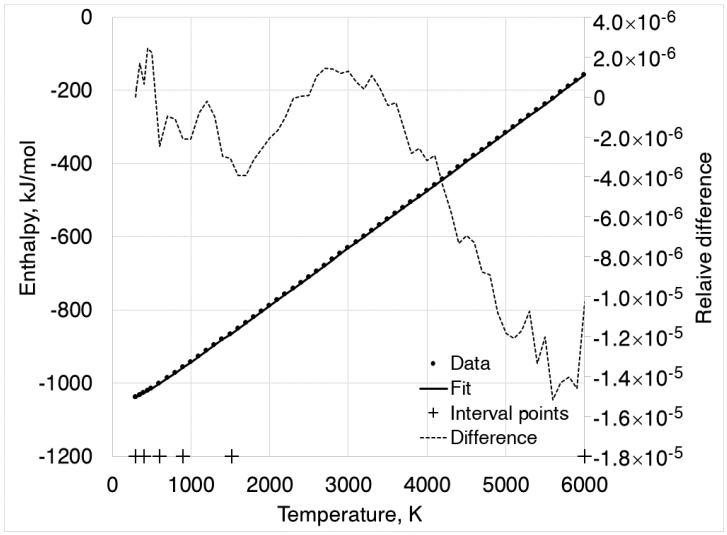


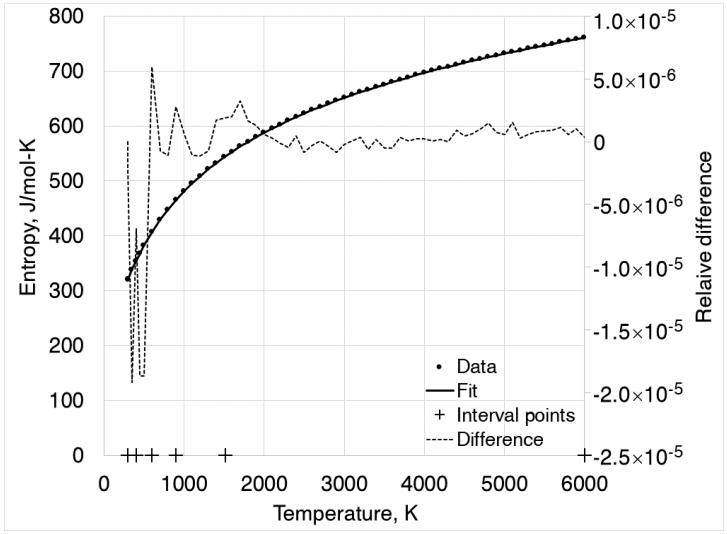


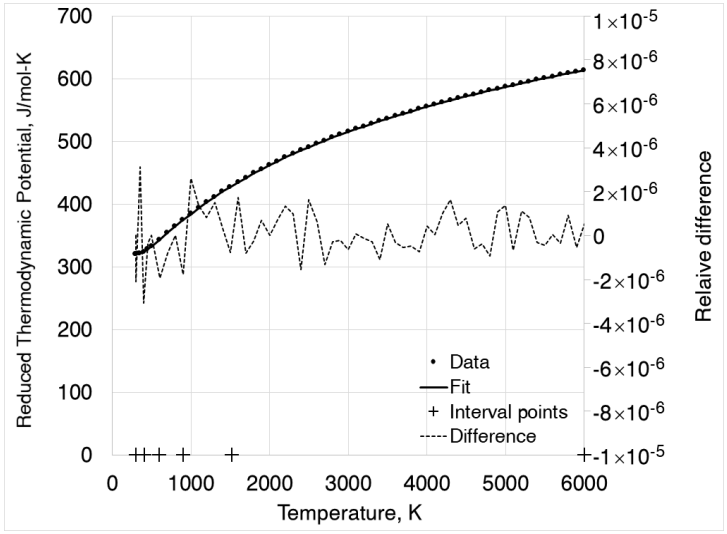
Sulfur Chloride Fluoride ($SClF_5$) – Cl1F5S1(g)

MW = 162.51 g/mol – $\Delta_f H^o$(298.15 K) = -1038.887 kJ/mol – S^o(298.15 K) = 319.931 J/mol-K – p^o = 0.1 MPa



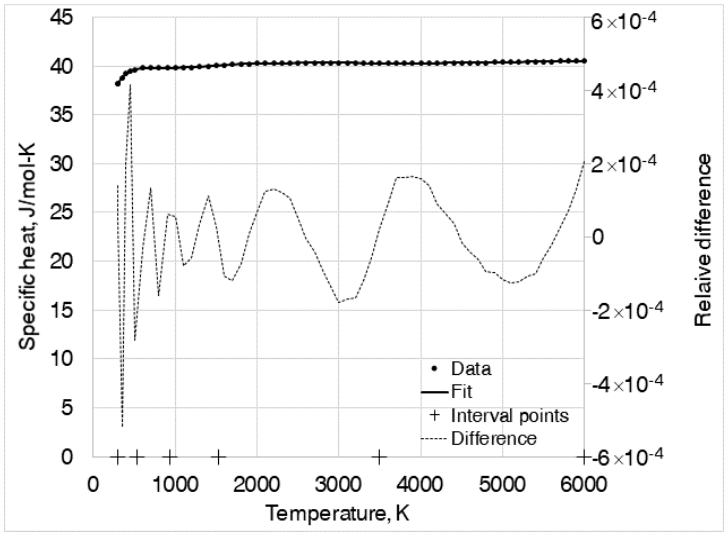


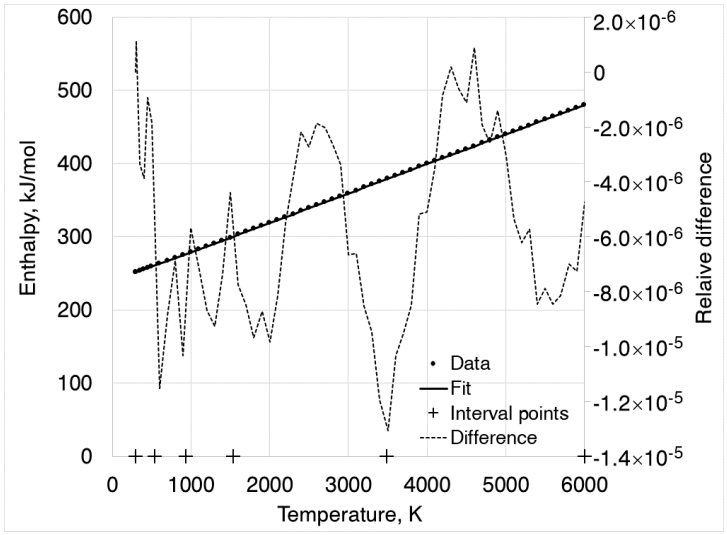


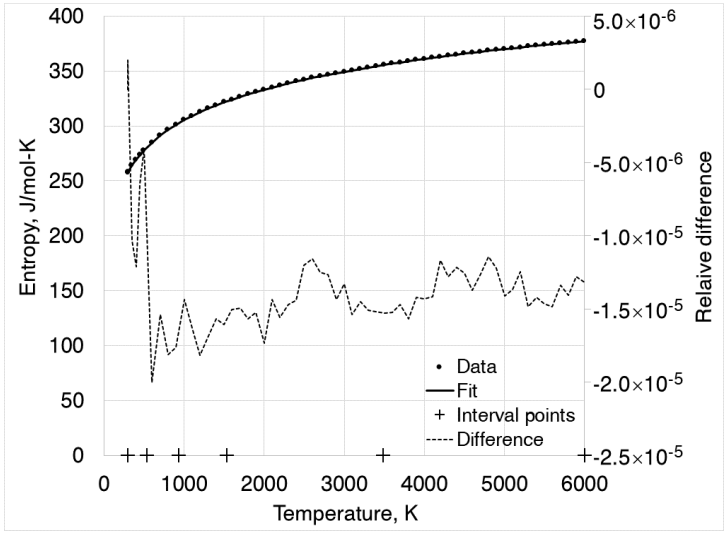


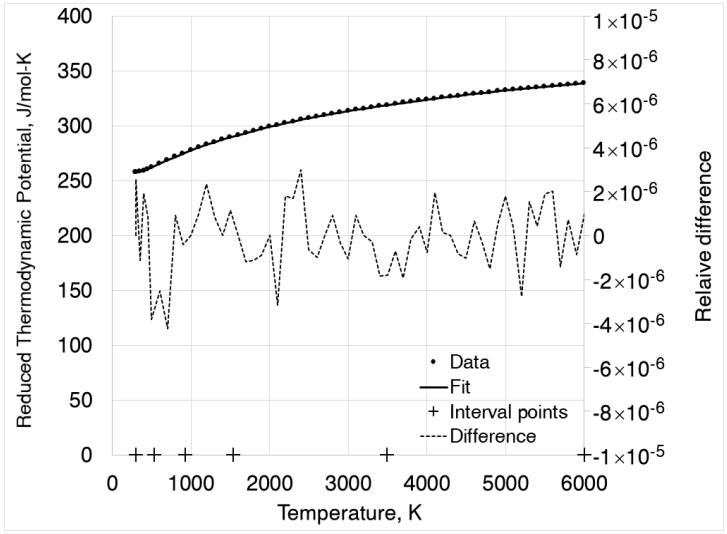
Iron Chloride (FeCl) – Cl1Fe1(g)

MW = 91.3 g/mol – $\Delta_f H^o$(298.15 K) = 251.04 kJ/mol – S^o(298.15 K) = 257.573 J/mol-K – p^o = 0.1 MPa



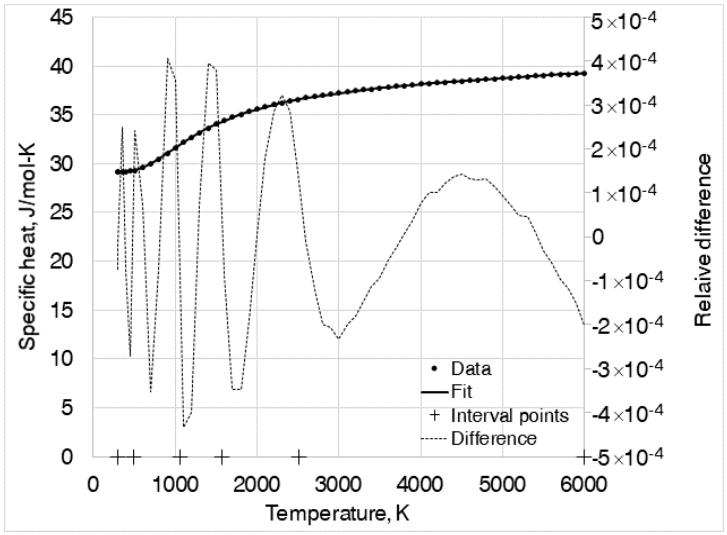


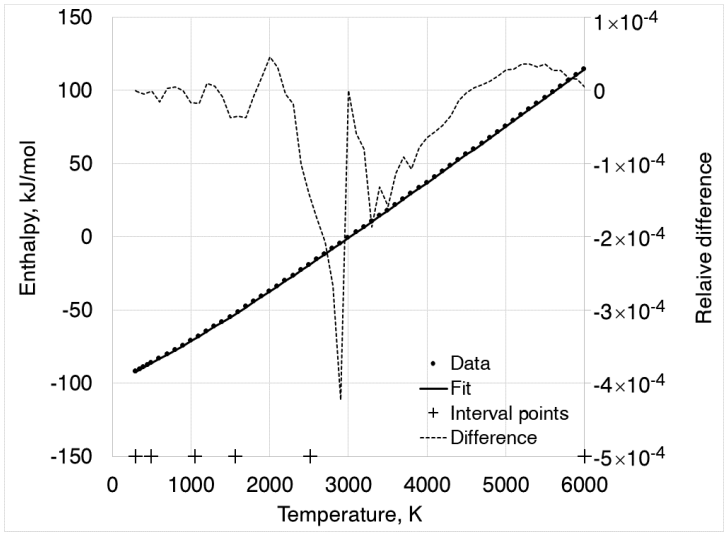


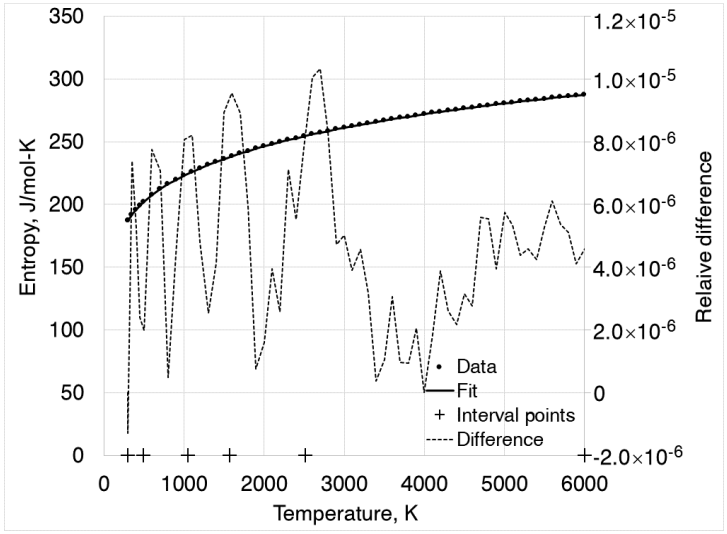


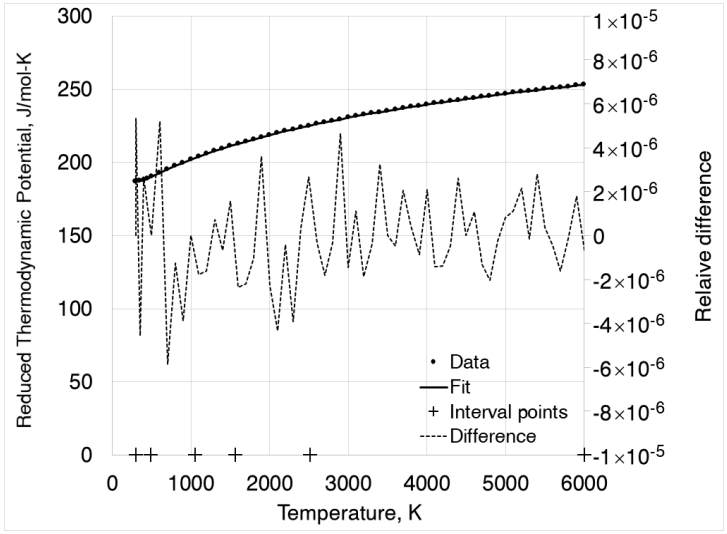
Hydrogen Chloride (HCl) – Cl1H1(g)

MW = 36.461 g/mol – $\Delta_f H^o$(298.15 K) = -92.312 kJ/mol – S^o(298.15 K) = 186.901 J/mol-K – p^o = 0.1 MPa



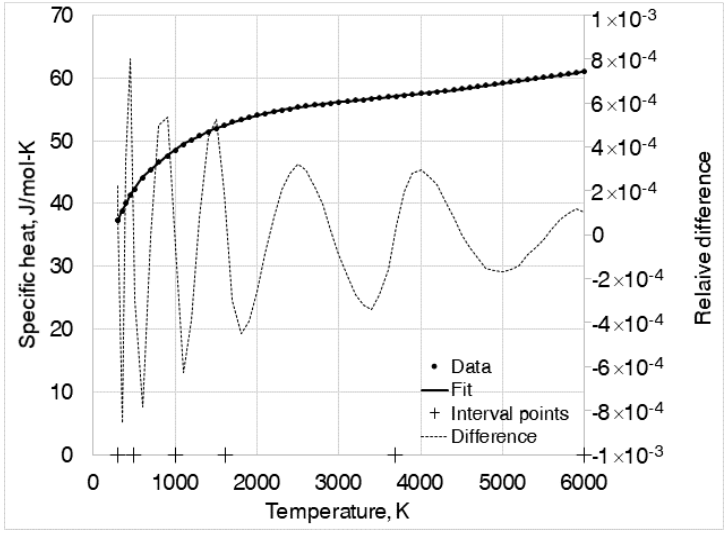


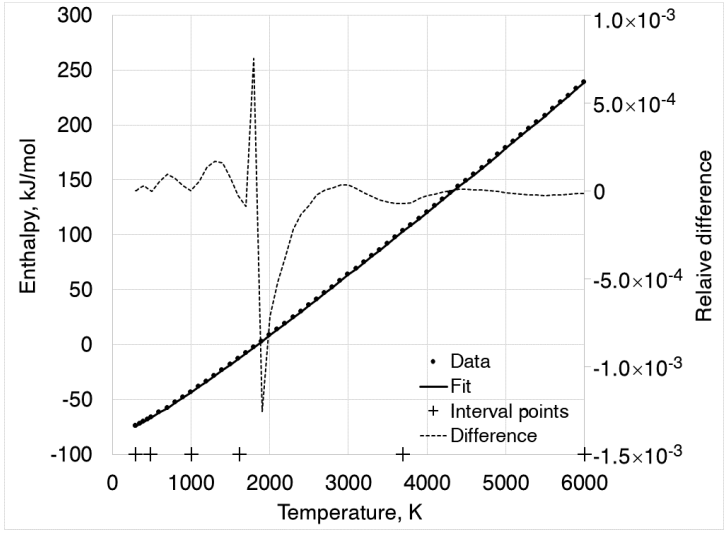


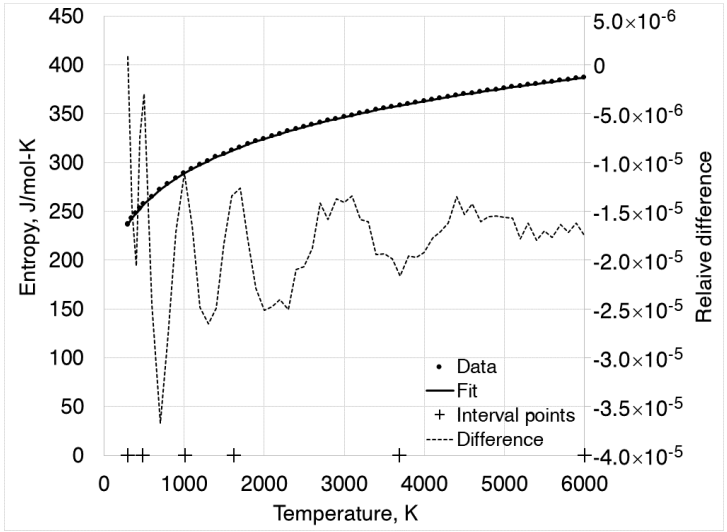


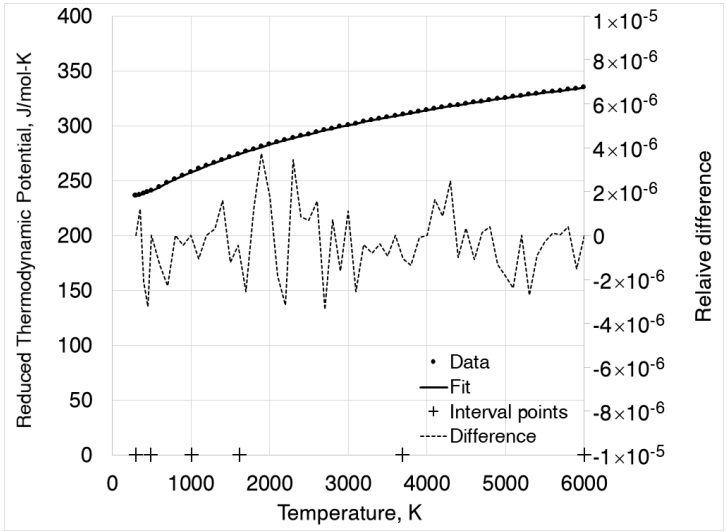
Hypochlorous Acid (HOCl) – Cl1H1O1(g)

MW = 52.46 g/mol – $\Delta_f H^o$(298.15 K) = -74.475 kJ/mol – S^o(298.15 K) = 236.504 J/mol-K – p^o = 0.1 MPa



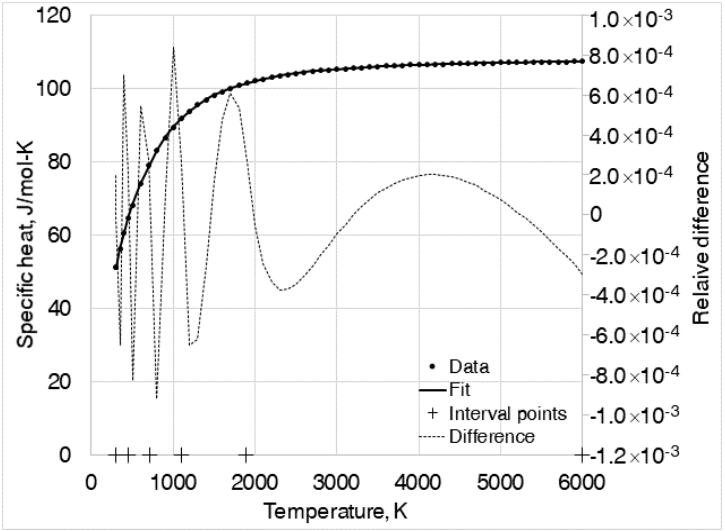


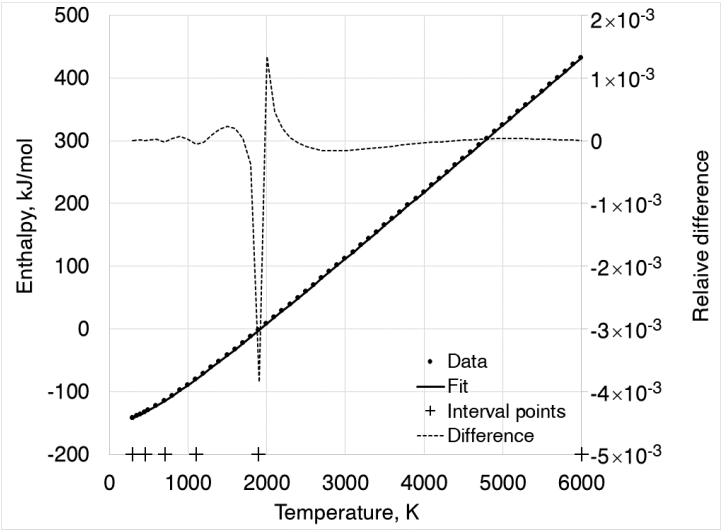


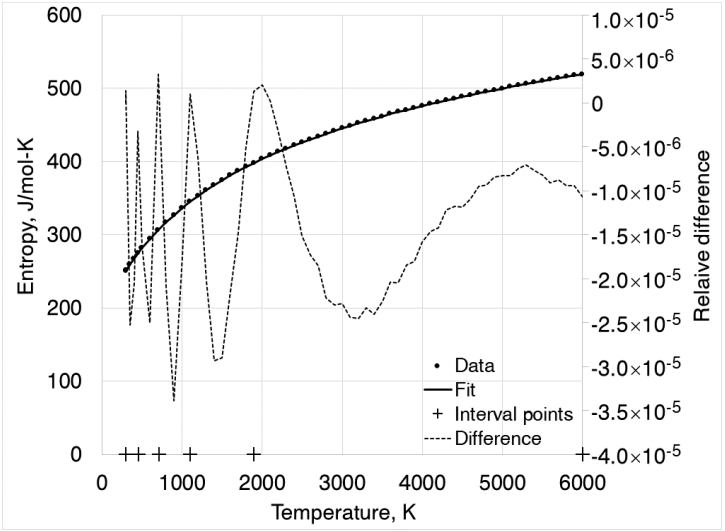


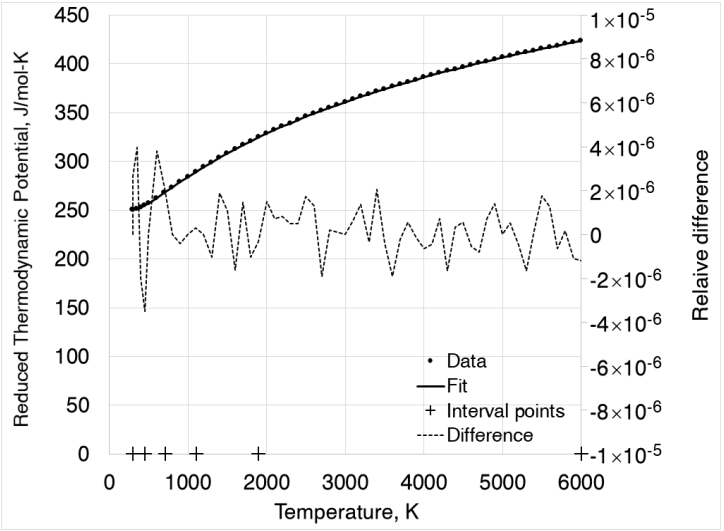
Chlorosilane (SiH_3Cl) – Cl1H3Si1(g)

MW = 66.562 g/mol – $\Delta_f H^o$(298.15 K) = -141.838 kJ/mol – S^o(298.15 K) = 250.761 J/mol-K – p^o = 0.1 MPa



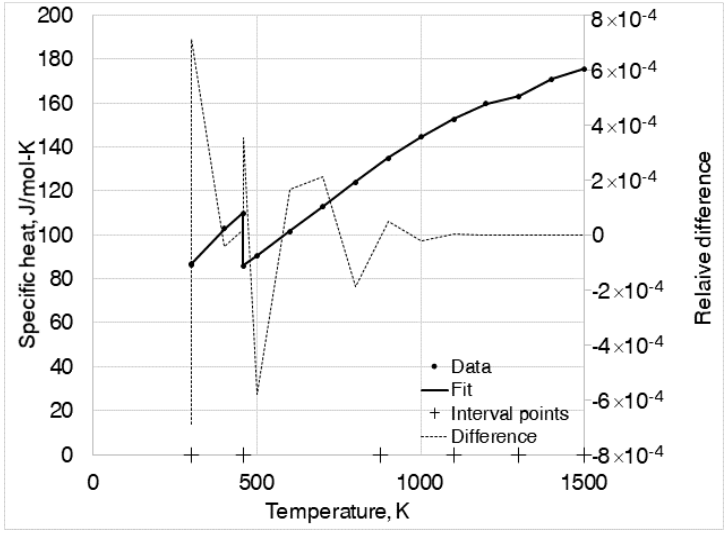


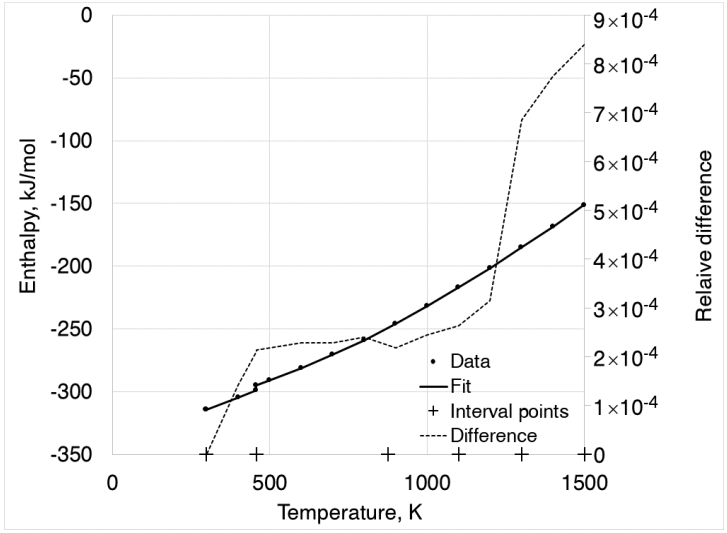


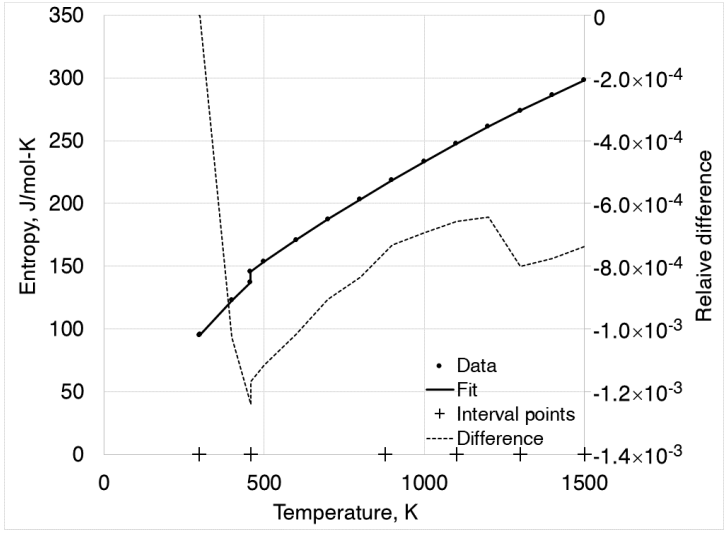


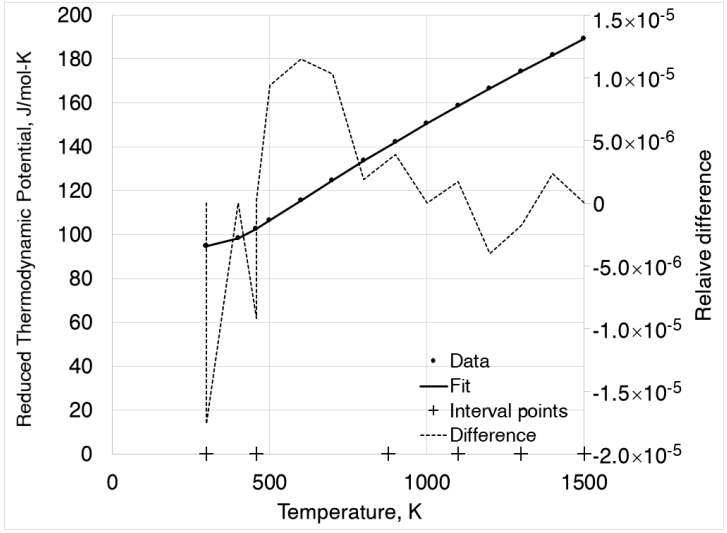
Ammonium Chloride (NH_4Cl) – Cl1H4N1(cr)

MW = 53.491 g/mol – $\Delta_f H^o$(298.15 K) = -314.553 kJ/mol – S^o(298.15 K) = 94.86 J/mol-K – p^o = 0.1 MPa



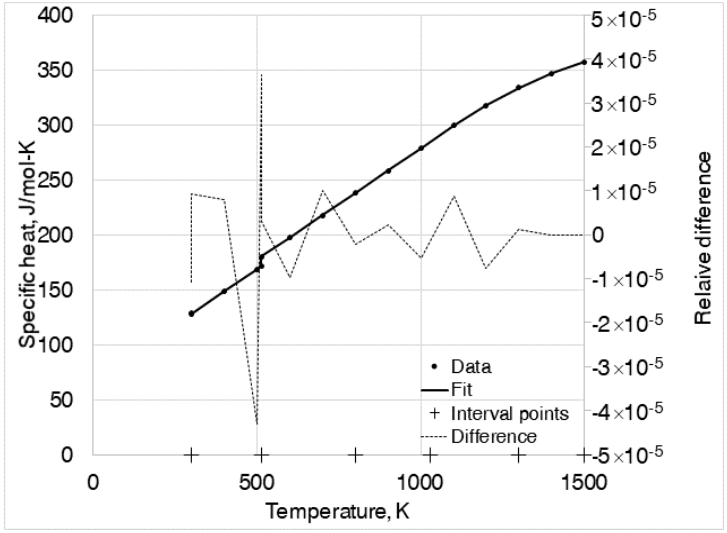


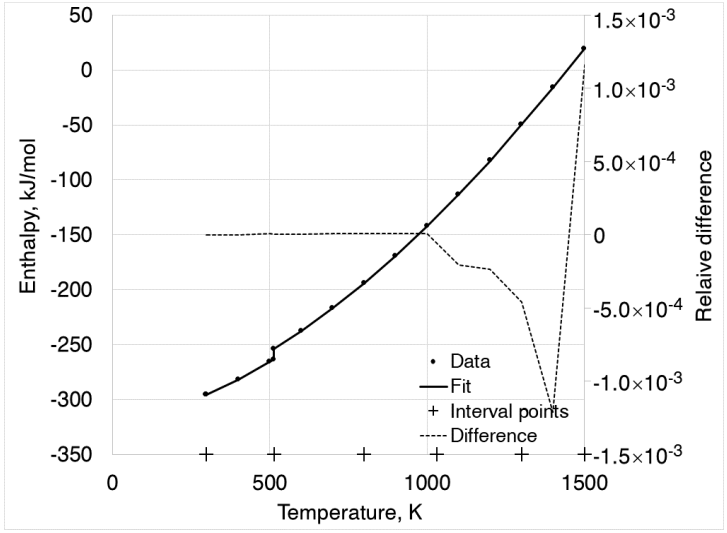


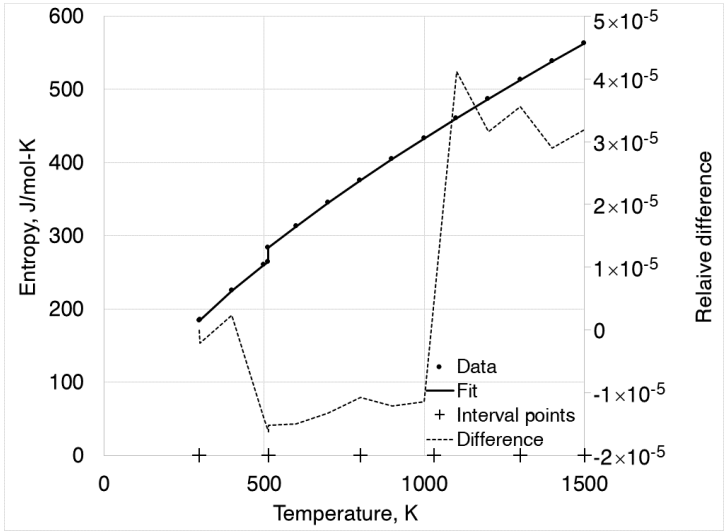


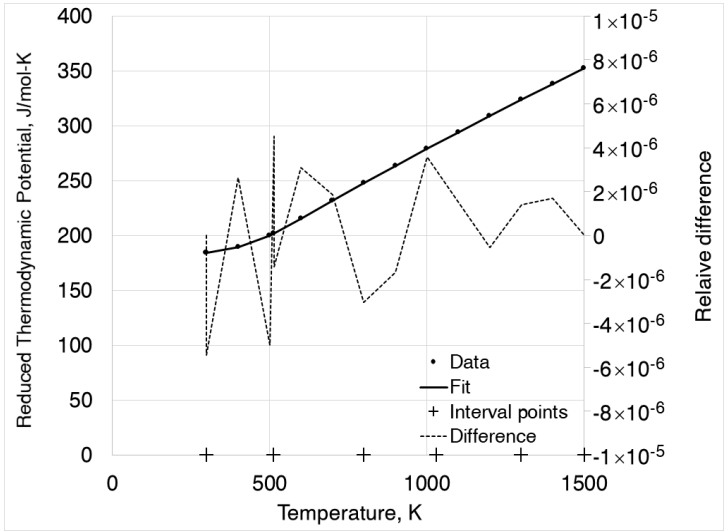
Ammonium Perchlorate (NH_4ClO_4) – Cl1H4N1O4(cr)

MW = 117.489 g/mol – $\Delta_f H^o$(298.15 K) = -295.767 kJ/mol – S^o(298.15 K) = 184.18 J/mol-K – p^o = 0.1 MPa



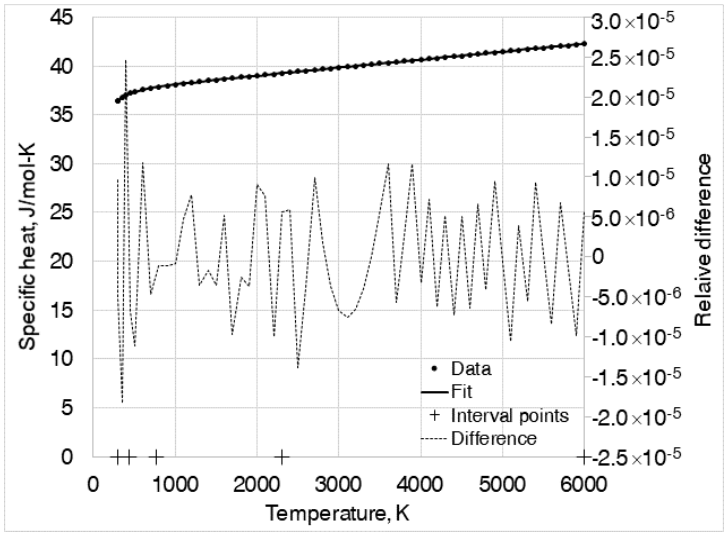


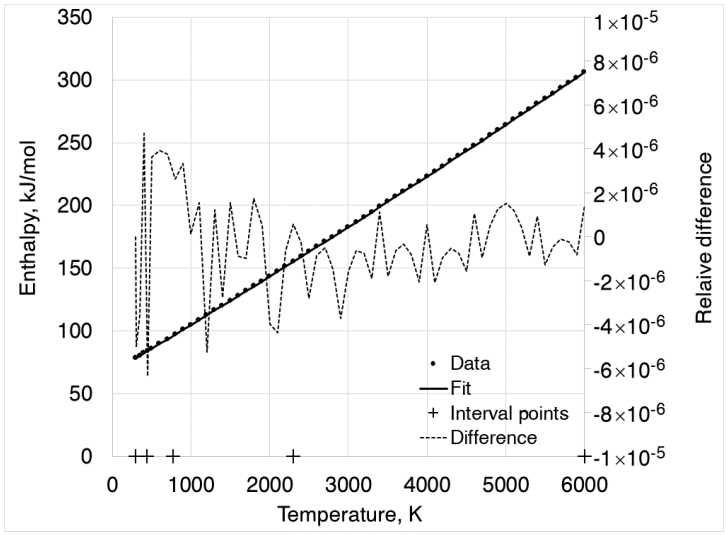


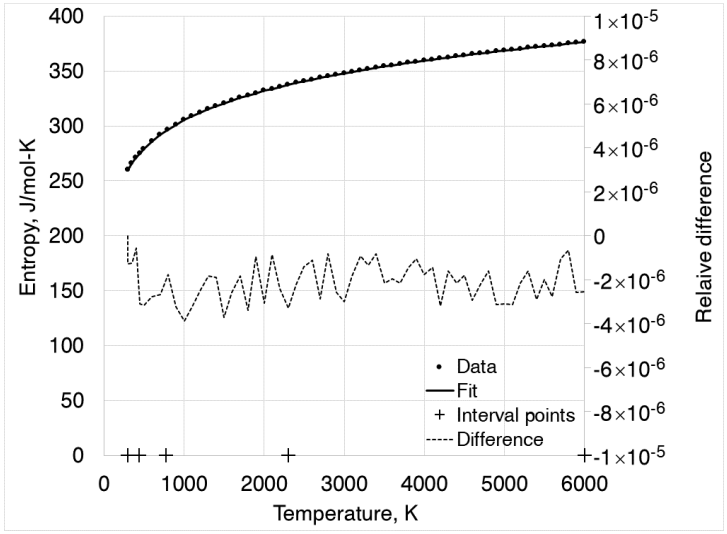


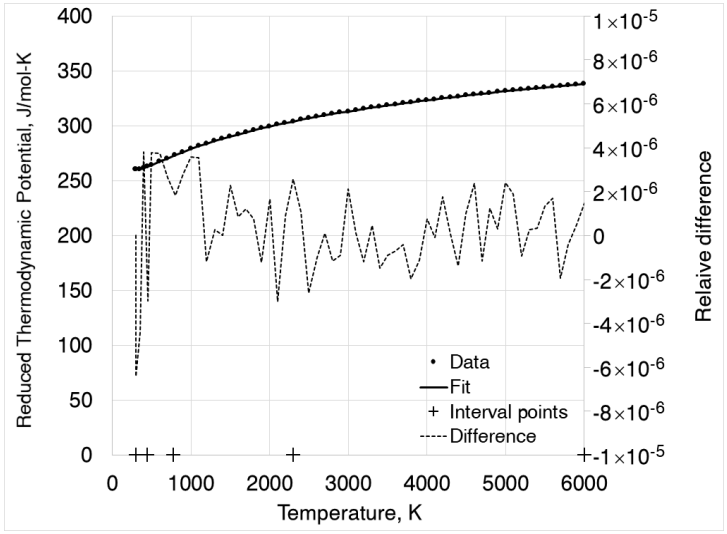
Mercury Chloride (HgCl) – Cl1Hg1(g)

MW = 236.04 g/mol – $\Delta_f H^o$(298.15 K) = 78.45 kJ/mol – S^o(298.15 K) = 260.005 J/mol-K – p^o = 0.1 MPa



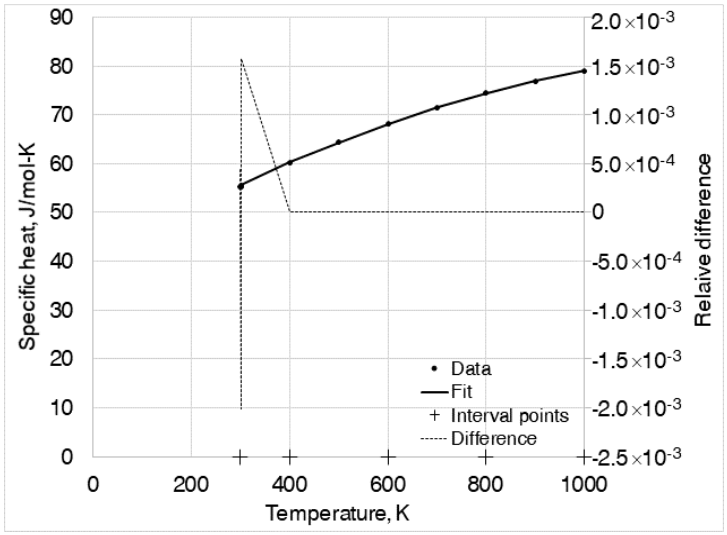


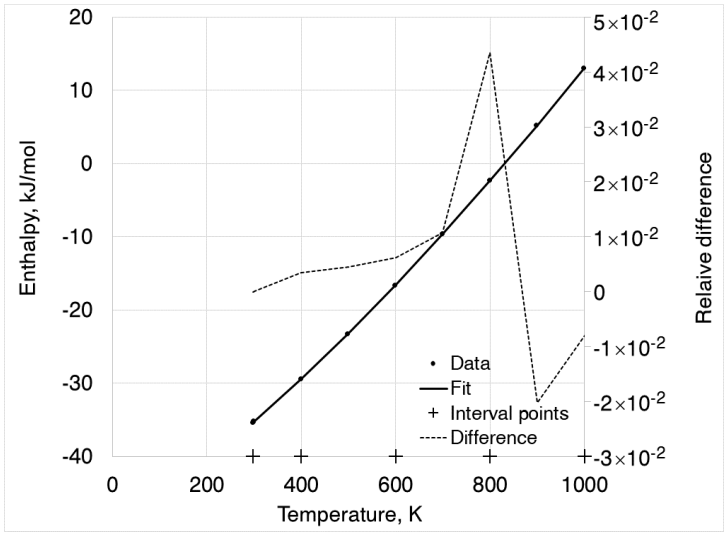


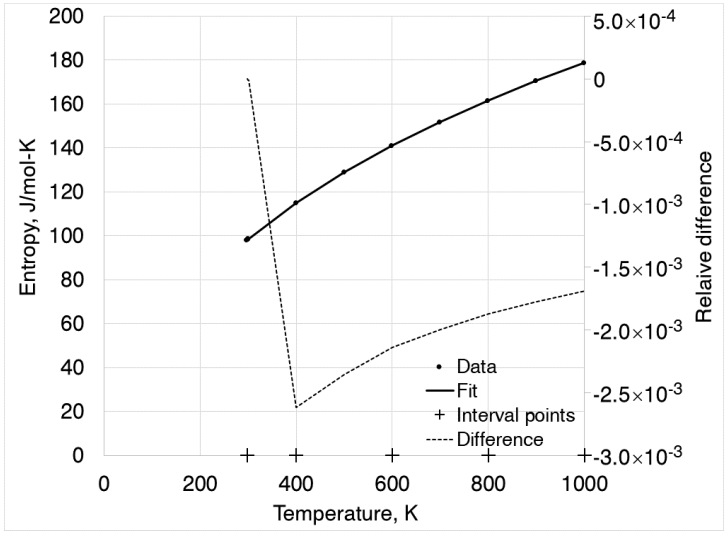


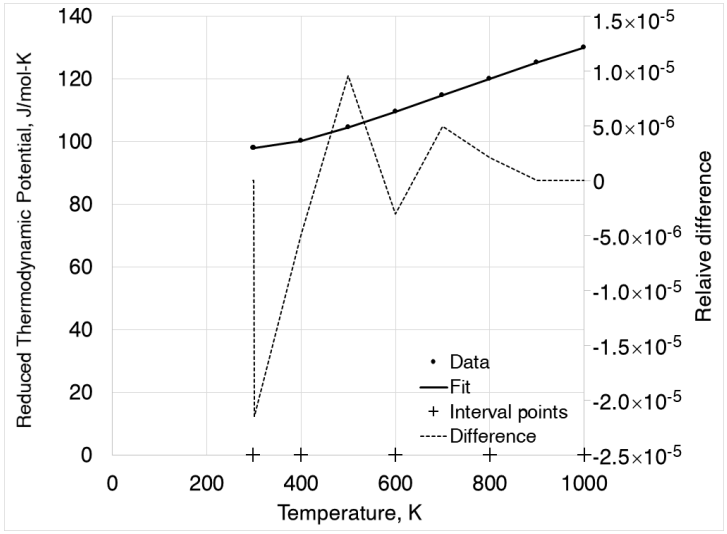
Iodine Chloride (ICl) – Cl1I1(cr)

MW = 162.358 g/mol – $\Delta_f H^o$(298.15 K) = -35.422 kJ/mol – S^o(298.15 K) = 97.927 J/mol-K – p^o = 0.1 MPa



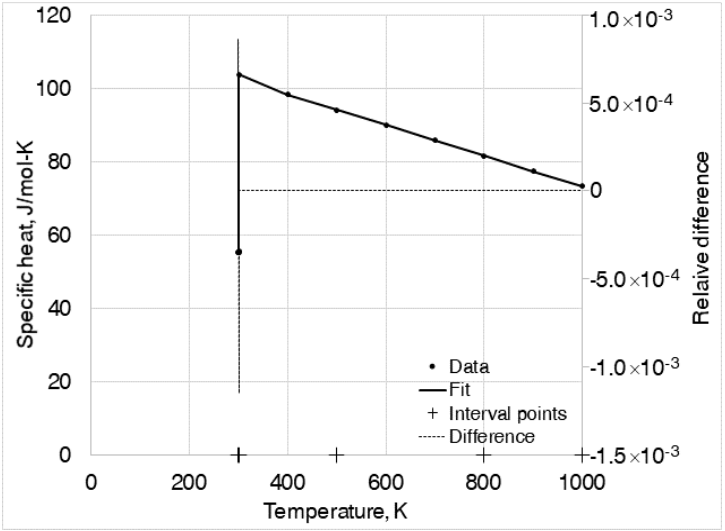


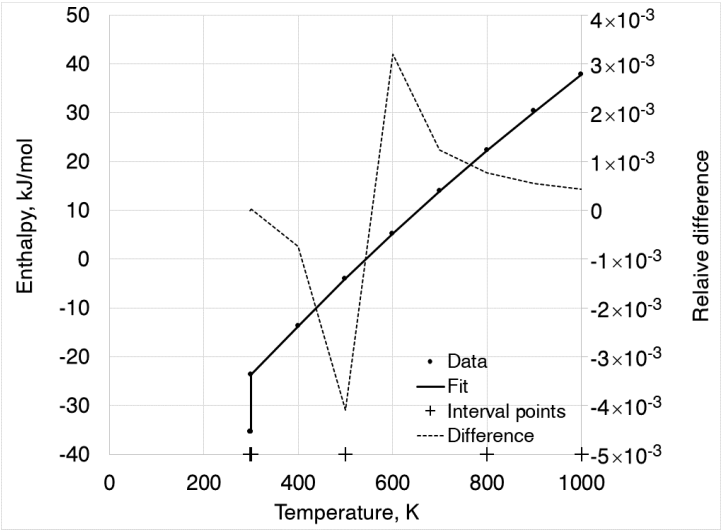


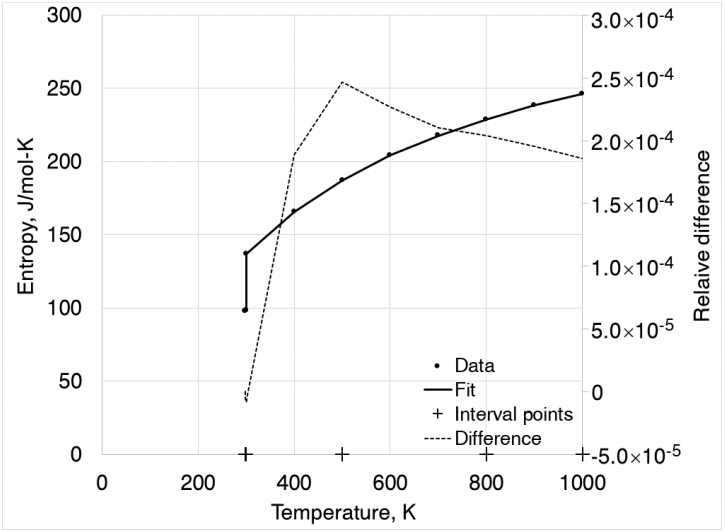


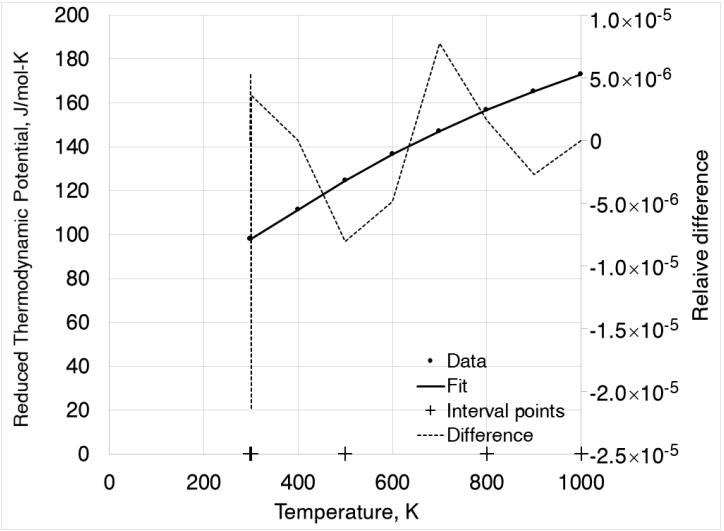
Iodine Chloride (ICl) – Cl1I1(cr,l)

MW = 162.358 g/mol – $\Delta_f H^o$(298.15 K) = -35.422 kJ/mol – S^o(298.15 K) = 97.927 J/mol-K – p^o = 0.1 MPa



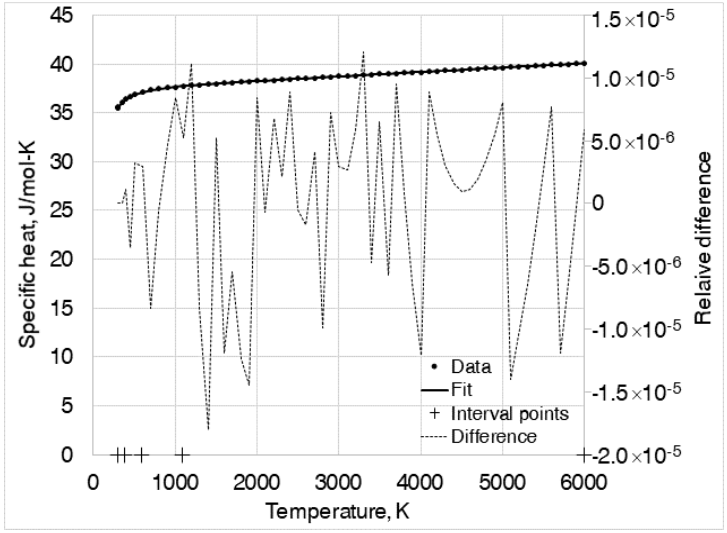


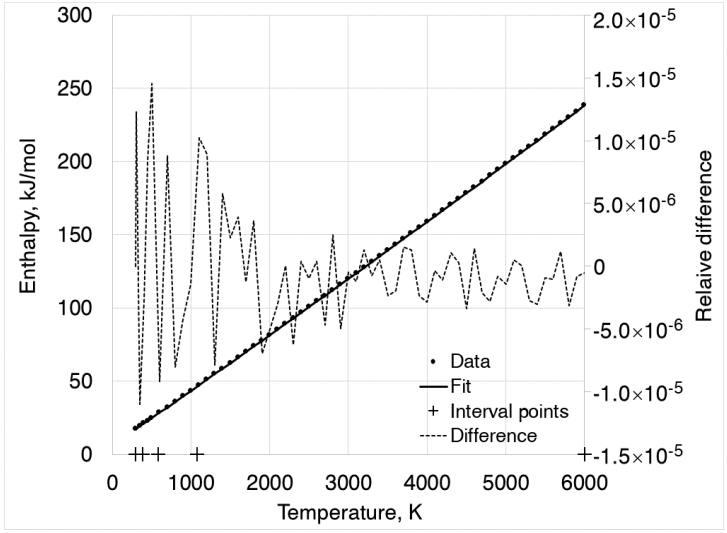


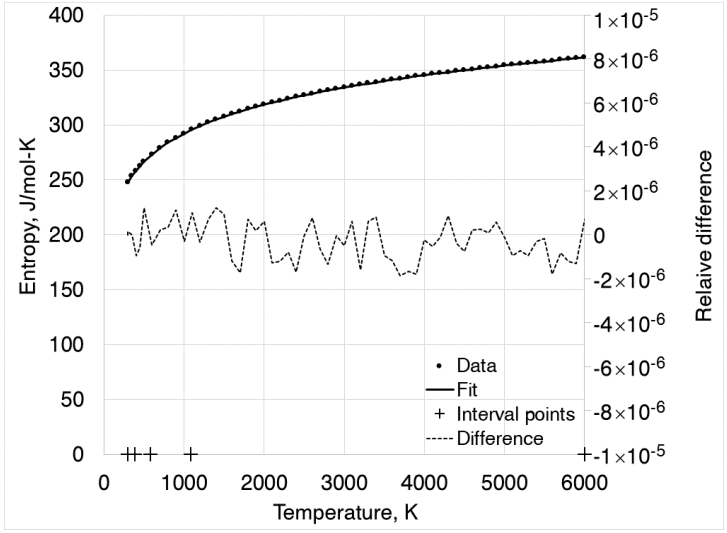


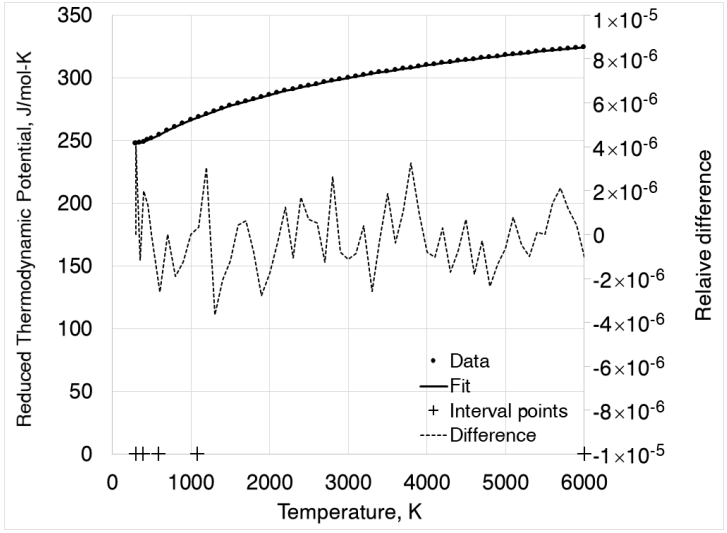
Iodine Chloride (ICl) – Cl1I1(g)

MW = 162.358 g/mol – $\Delta_f H^o$(298.15 K) = 17.506 kJ/mol – S^o(298.15 K) = 247.567 J/mol-K – p^o = 0.1 MPa



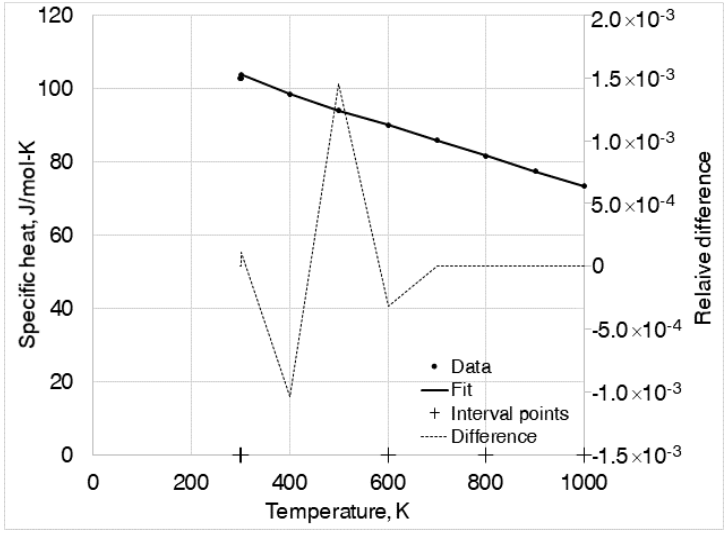


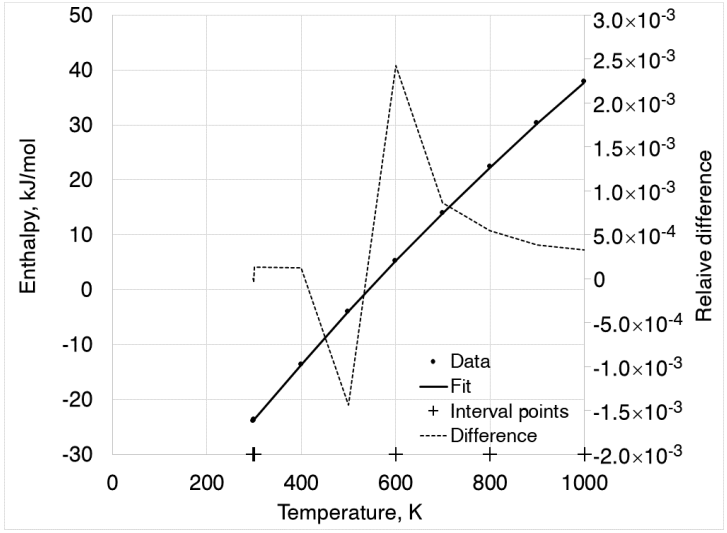


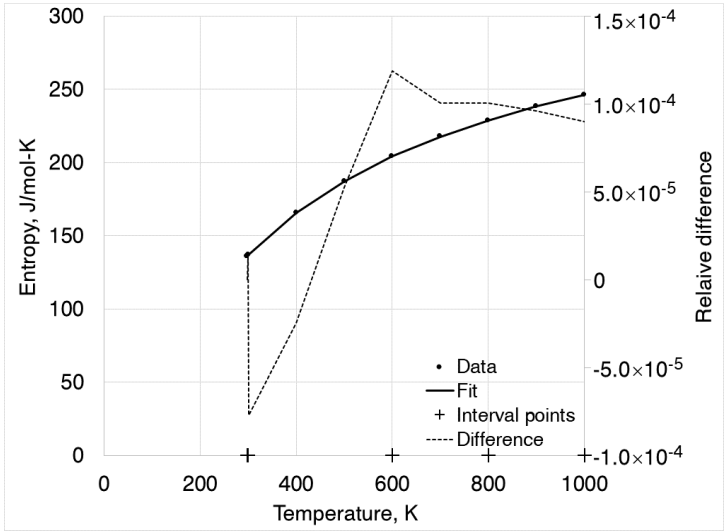


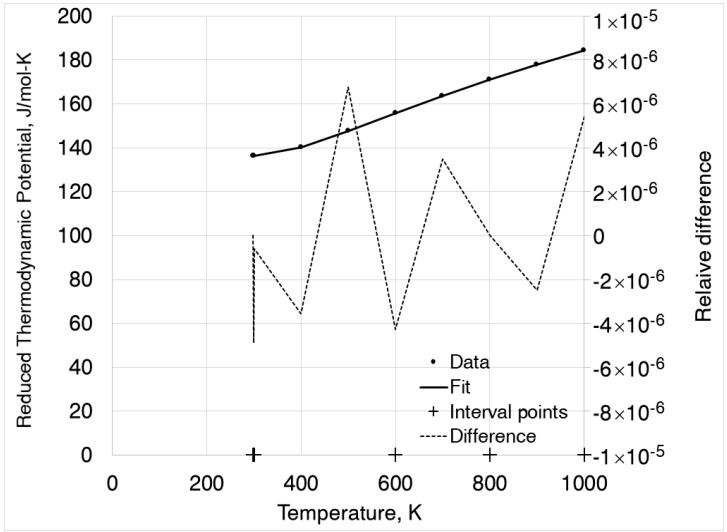
Iodine Chloride (ICl) – Cl1I1(l)

MW = 162.358 g/mol – $\Delta_f H^o$(298.15 K) = -23.929 kJ/mol – S^o(298.15 K) = 136.167 J/mol-K – p^o = 0.1 MPa



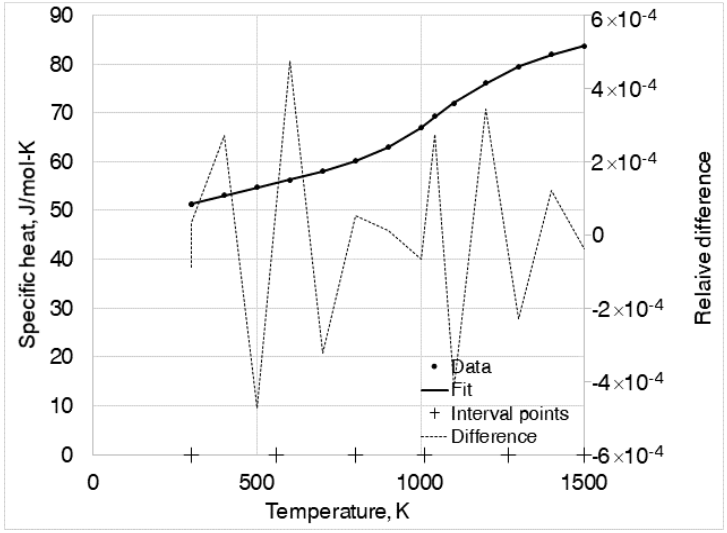


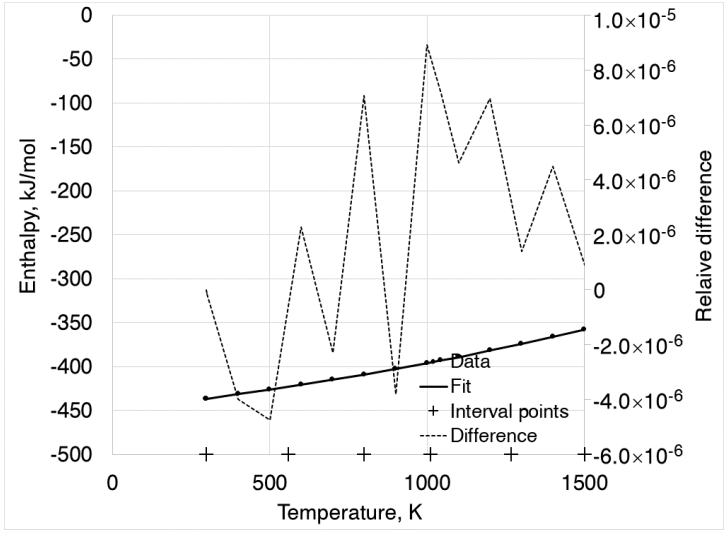


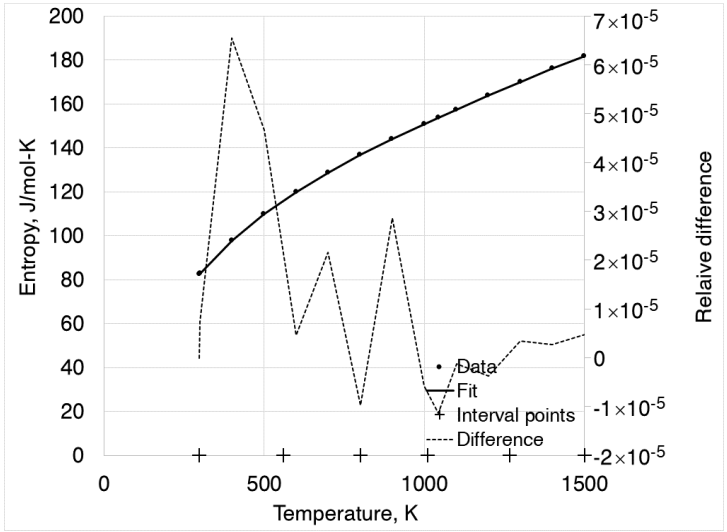


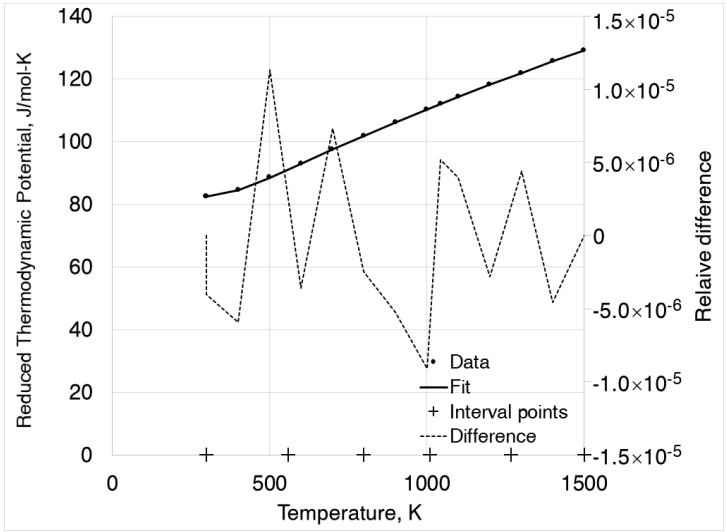
Potassium Chloride (KCl) – Cl1K1(cr)

MW = 74.551 g/mol – $\Delta_f H^o$(298.15 K) = -436.684 kJ/mol – S^o(298.15 K) = 82.554 J/mol-K – p^o = 0.1 MPa



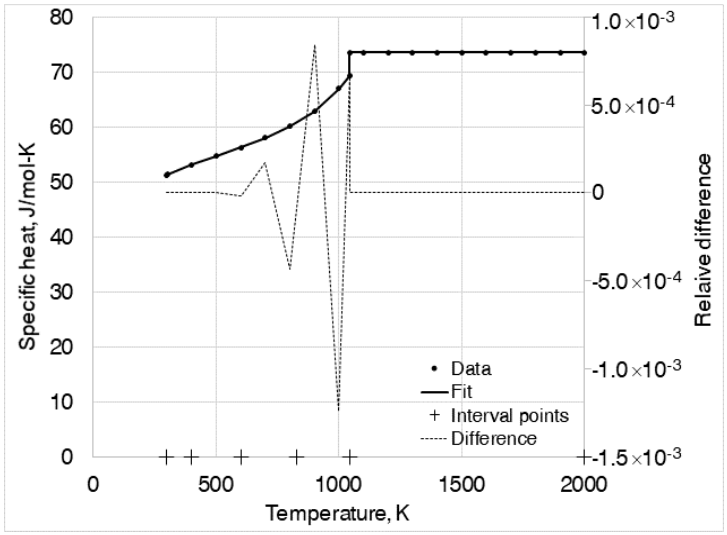


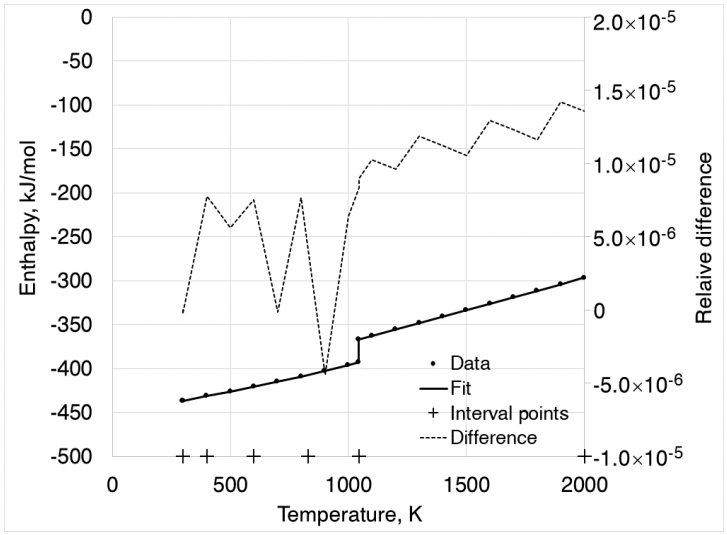


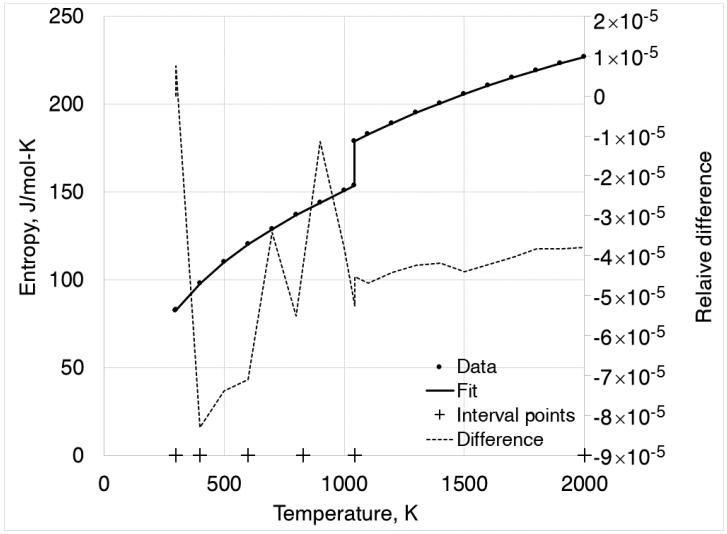


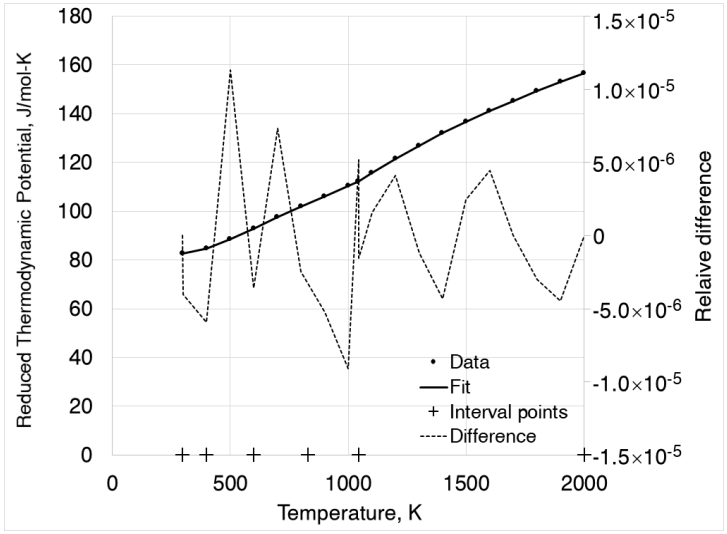
Potassium Chloride (KCl) – Cl1K1(cr,l)

MW = 74.551 g/mol – $\Delta_f H^o$(298.15 K) = -436.684 kJ/mol – S^o(298.15 K) = 82.554 J/mol-K – p^o = 0.1 MPa



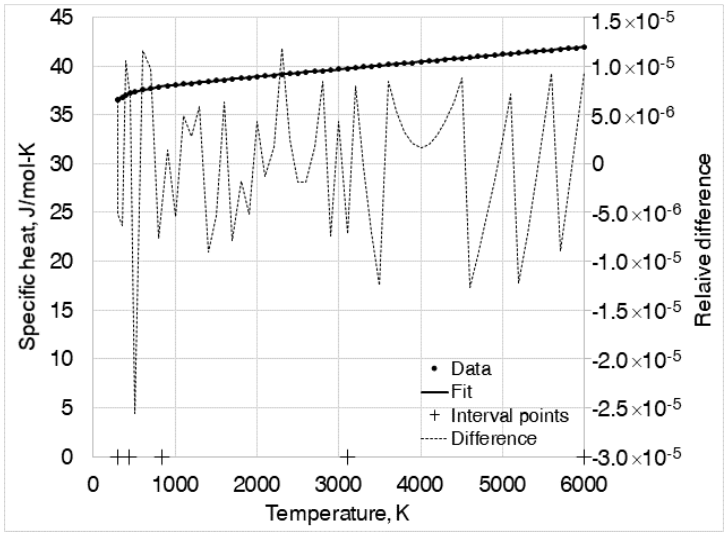


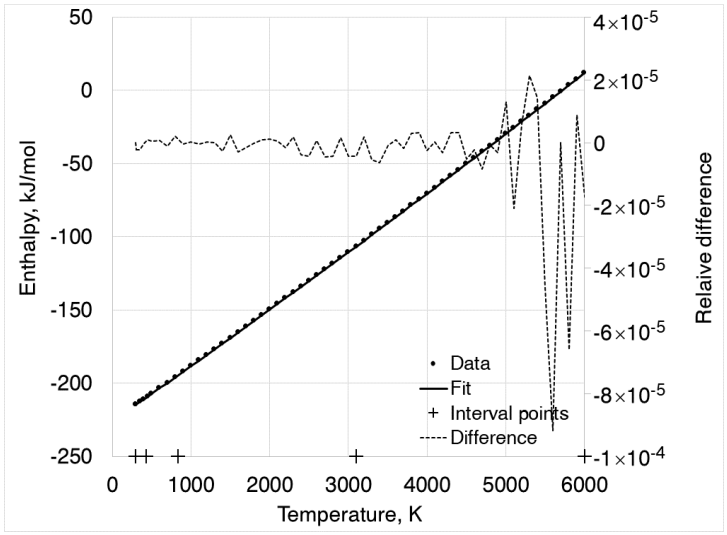


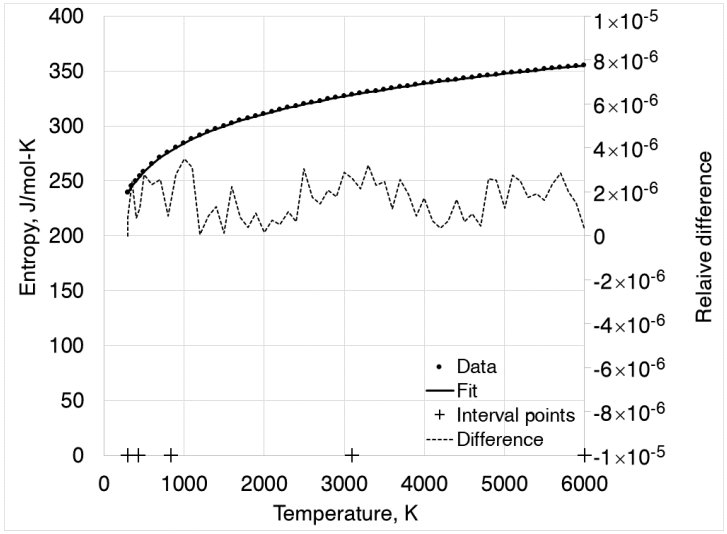


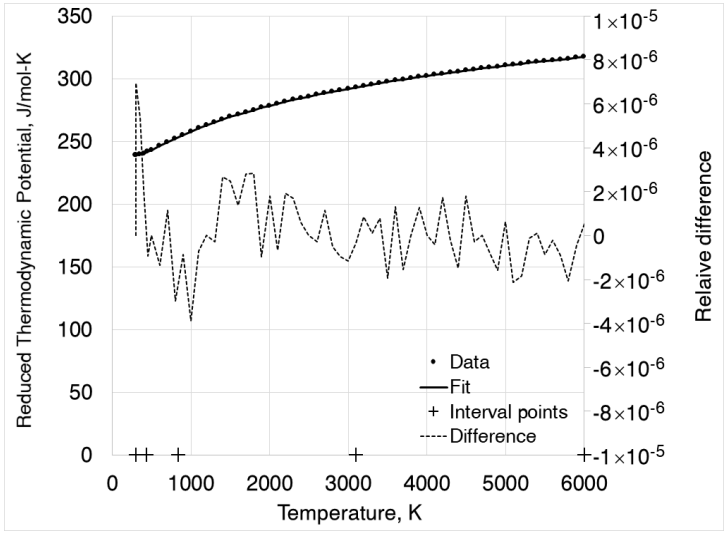
Potassium Chloride (KCl) – Cl1K1(g)

MW = 74.551 g/mol – $\Delta_f H^o$(298.15 K) = -214.681 kJ/mol – S^o(298.15 K) = 239.087 J/mol-K – p^o = 0.1 MPa



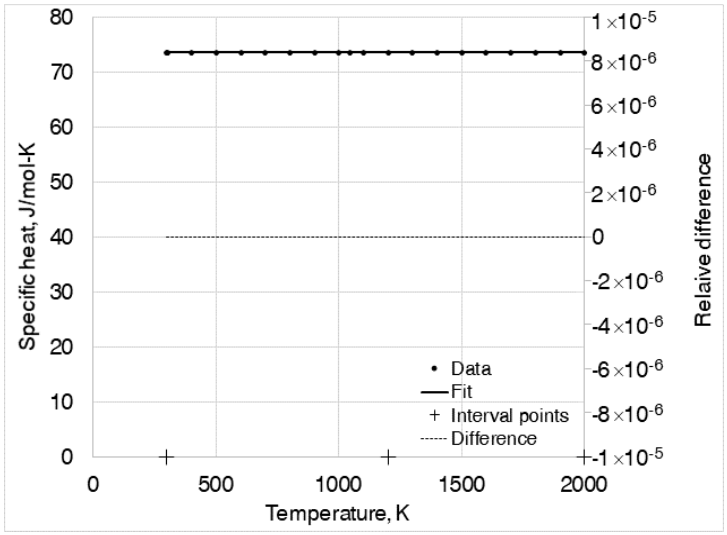


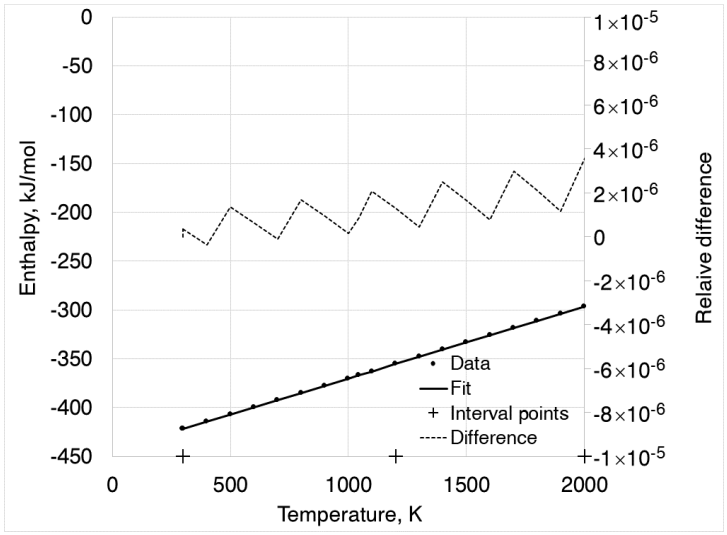


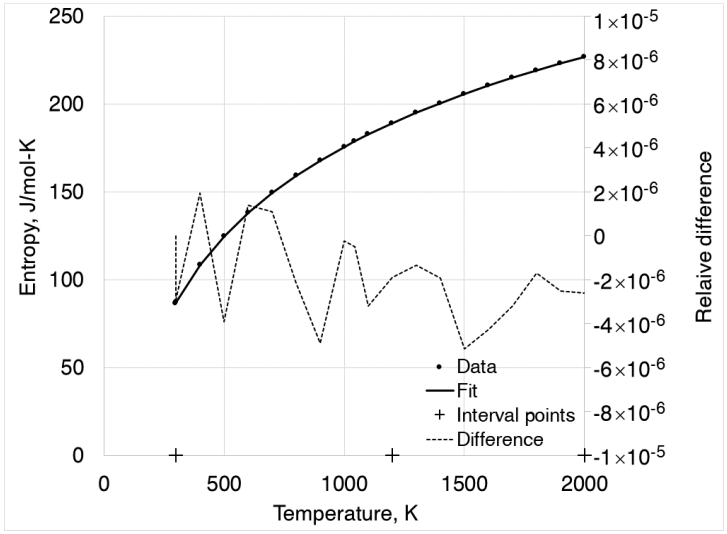


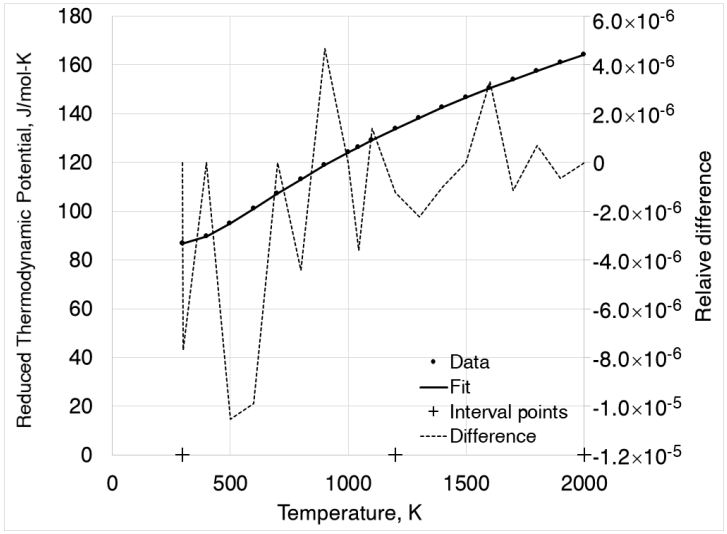
Potassium Chloride (KCl) – Cl1K1(l)

MW = 74.551 g/mol – $\Delta_f H^o$(298.15 K) = -421.791 kJ/mol – S^o(298.15 K) = 86.657 J/mol-K – p^o = 0.1 MPa



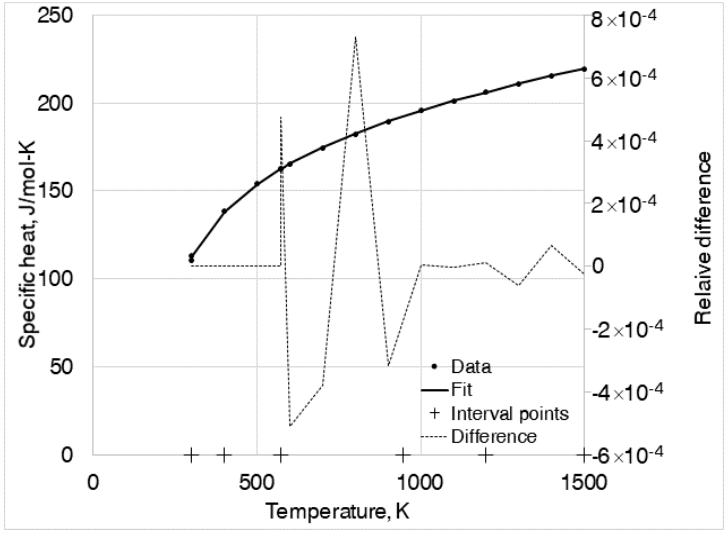


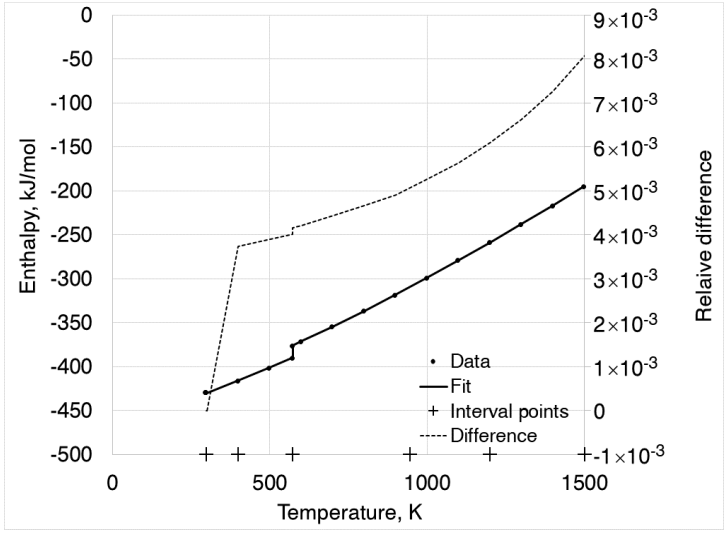


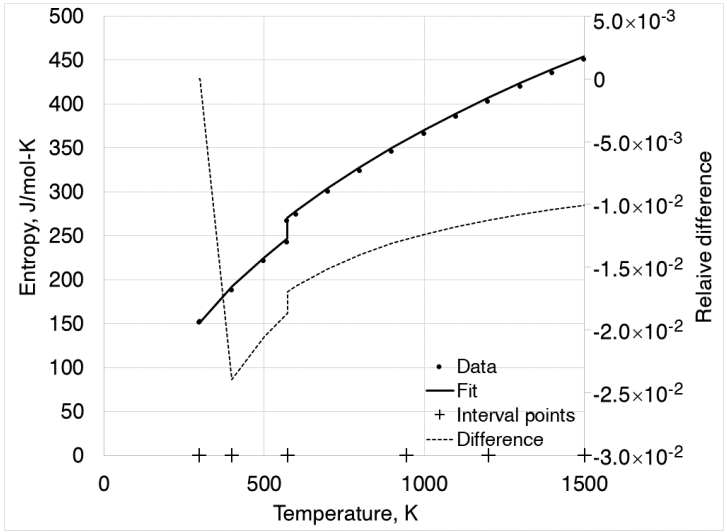


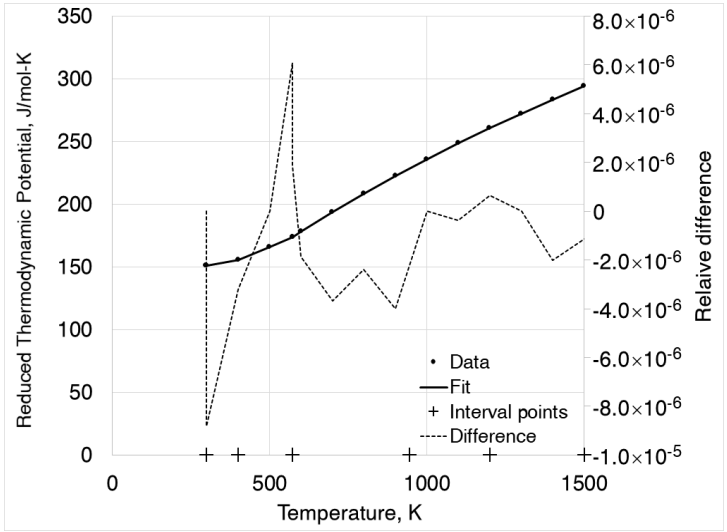
Potassium Perchlorate ($KClO_4$) – Cl1K1O4(cr)

MW = 138.549 g/mol – $\Delta_f H^o$(298.15 K) = -430.115 kJ/mol – S^o(298.15 K) = 151.042 J/mol-K – p^o = 0.1 MPa



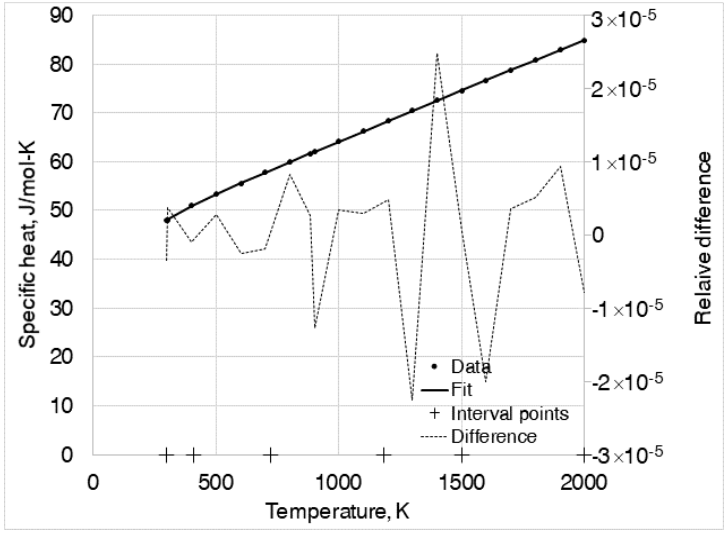


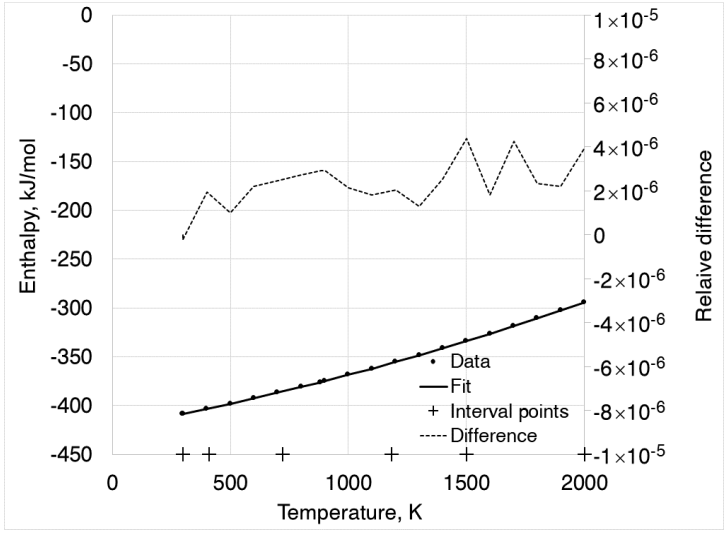


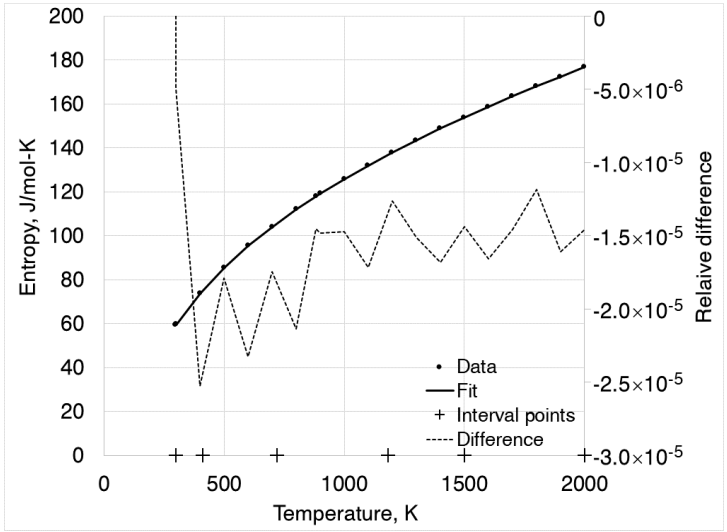


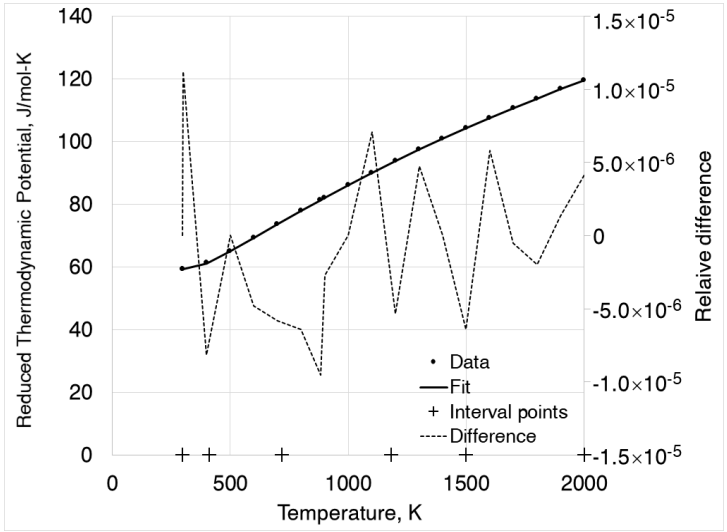
Lithium Chloride (LiCl) – Cl1Li1(cr)

MW = 42.394 g/mol – $\Delta_f H^o$(298.15 K) = -408.266 kJ/mol – S^o(298.15 K) = 59.3 J/mol-K – p^o = 0.1 MPa



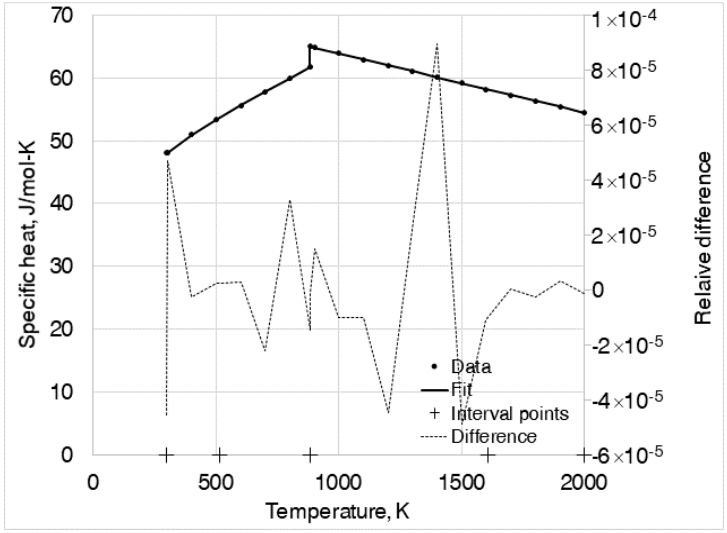


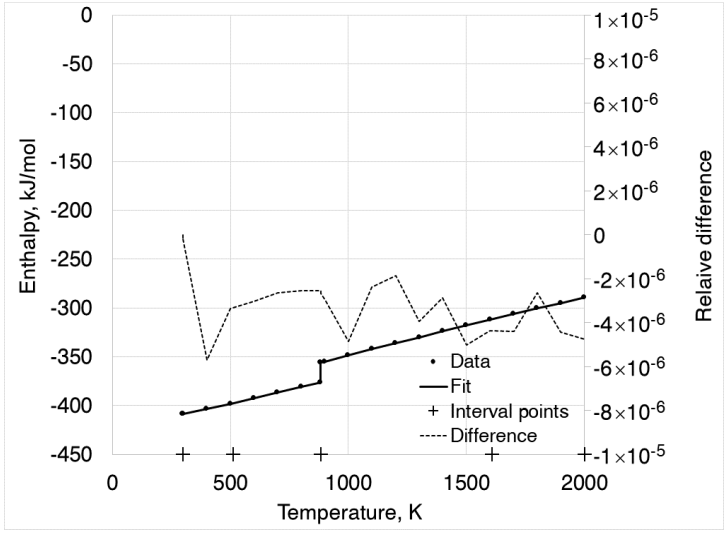


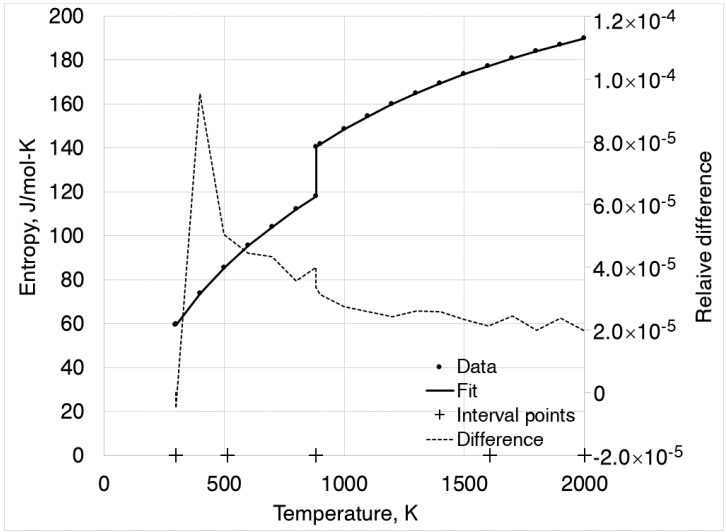


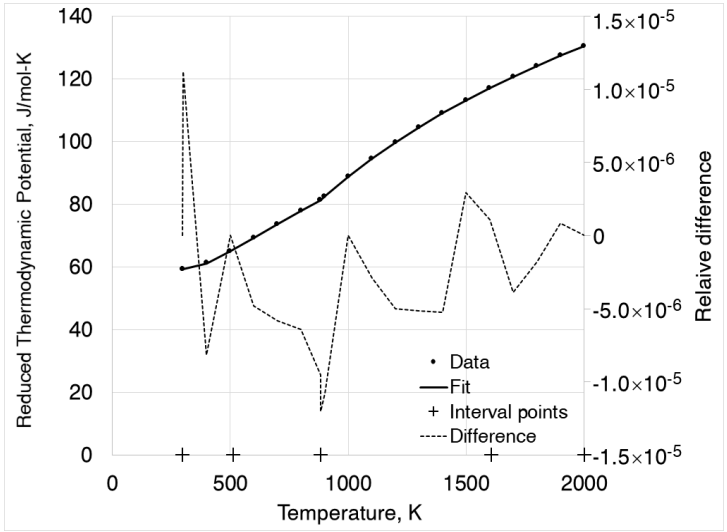
Lithium Chloride (LiCl) – Cl1Li1(cr,l)

MW = 42.394 g/mol – $\Delta_f H^o$(298.15 K) = -408.266 kJ/mol – S^o(298.15 K) = 59.3 J/mol-K – p^o = 0.1 MPa



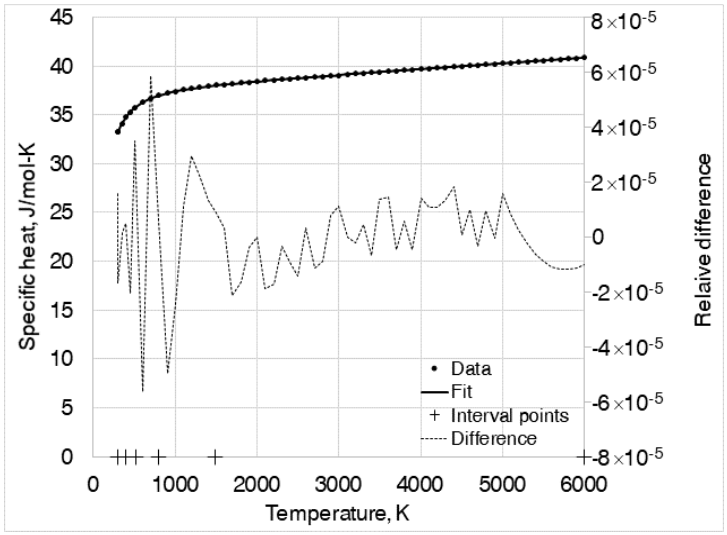


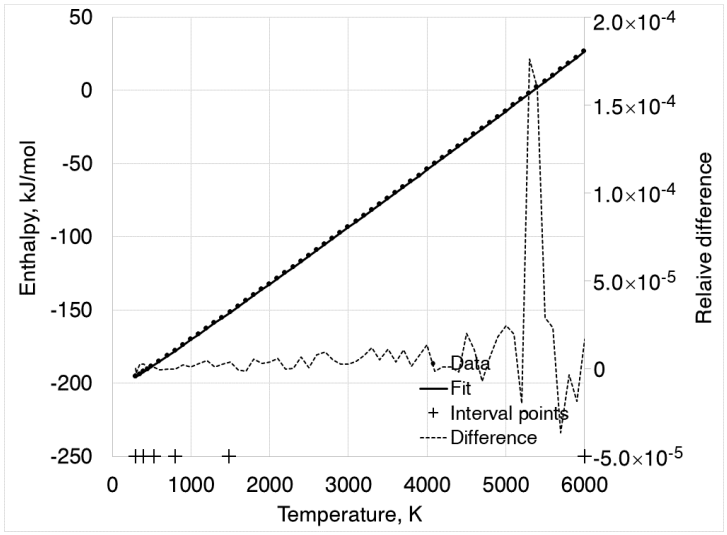


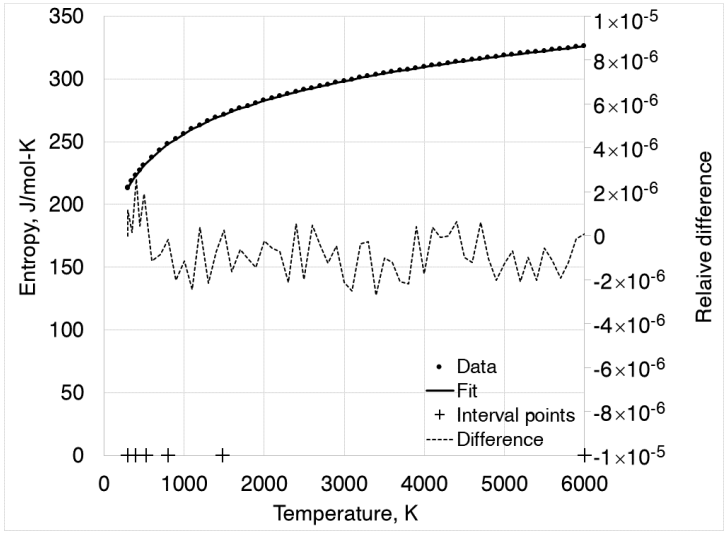


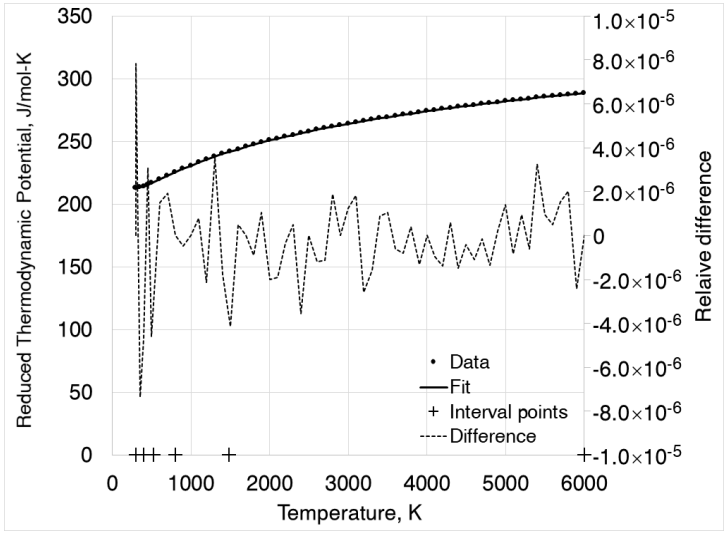
Lithium Chloride (LiCl) – Cl1Li1(g)

MW = 42.394 g/mol – $\Delta_f H^o$(298.15 K) = -195.719 kJ/mol – S^o(298.15 K) = 212.919 J/mol-K – p^o = 0.1 MPa



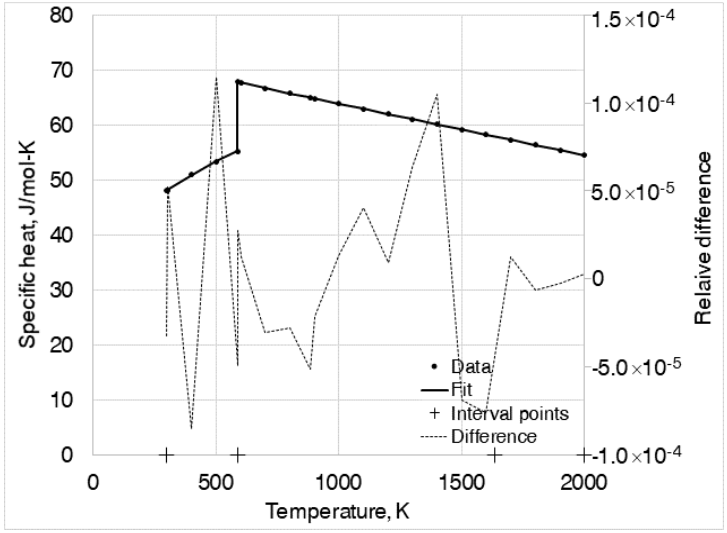


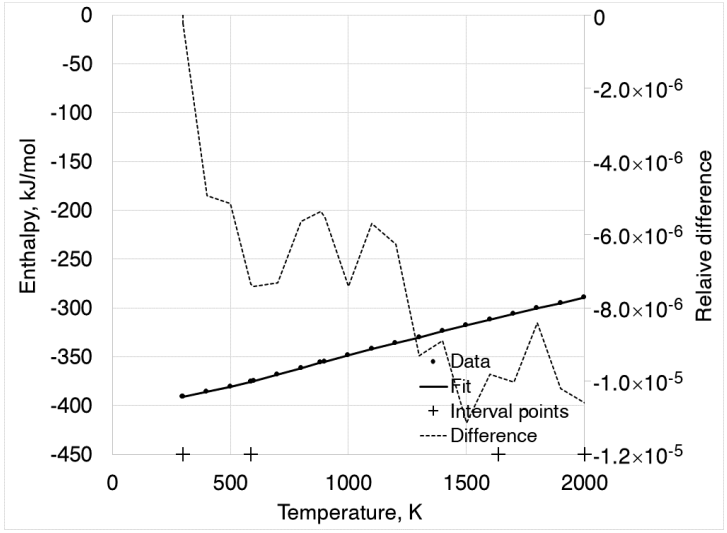


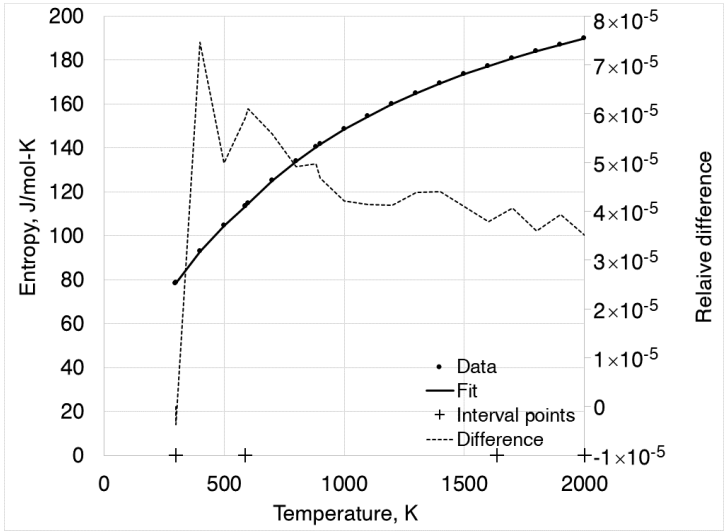


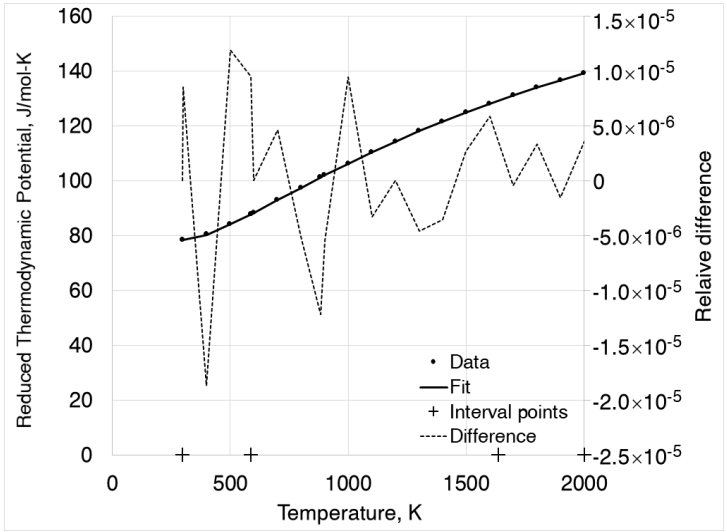
Lithium Chloride (LiCl) – Cl1Li1(l)

MW = 42.394 g/mol – $\Delta_f H^o$(298.15 K) = -390.755 kJ/mol – S^o(298.15 K) = 78.435 J/mol-K – p^o = 0.1 MPa



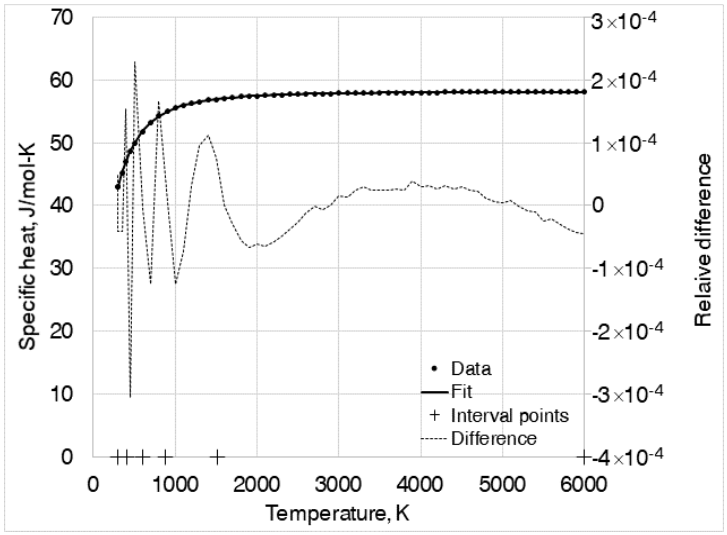


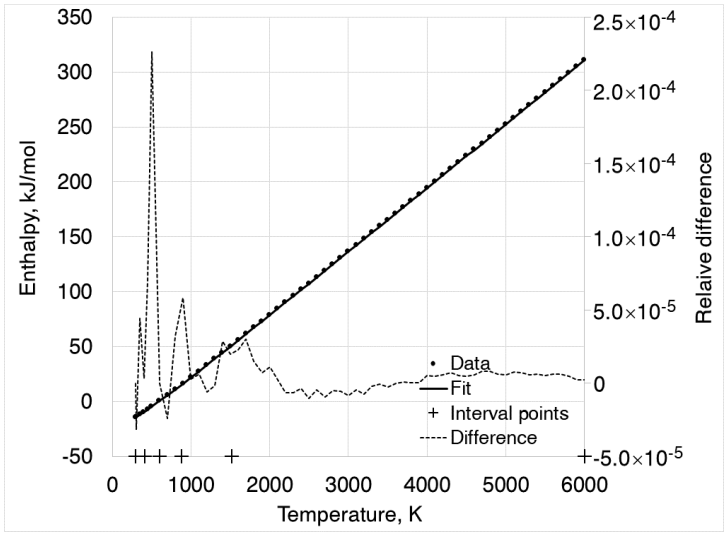


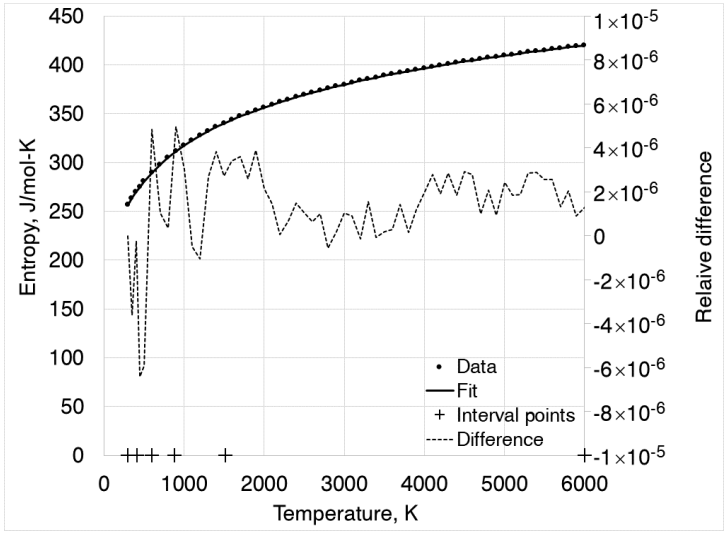


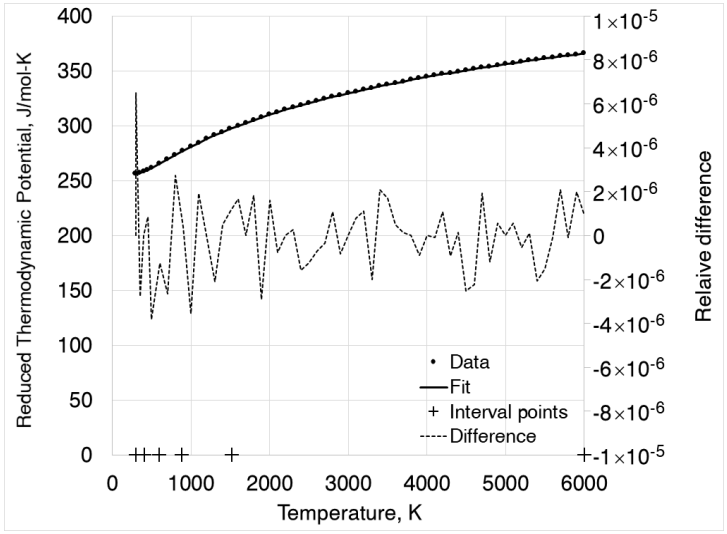
Lithium Hypochlorite (LiOCl) – Cl1Li1O1(g)

MW = 58.393 g/mol – $\Delta_f H^o$(298.15 K) = -14.226 kJ/mol – S^o(298.15 K) = 256.417 J/mol-K – p^o = 0.1 MPa



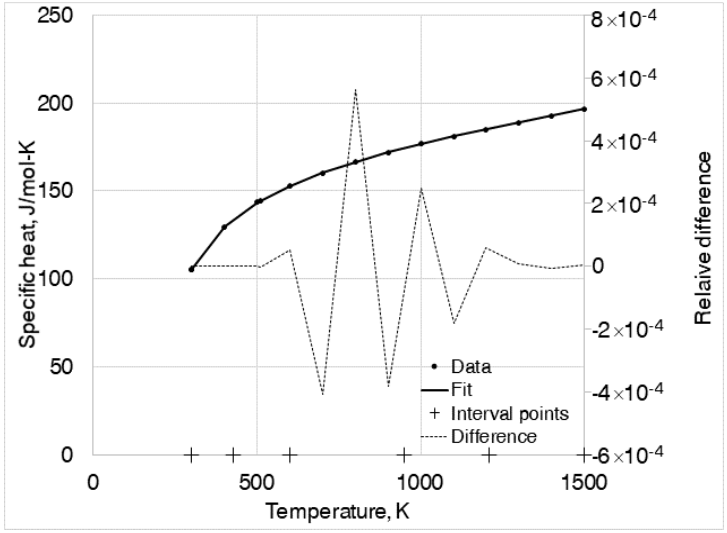


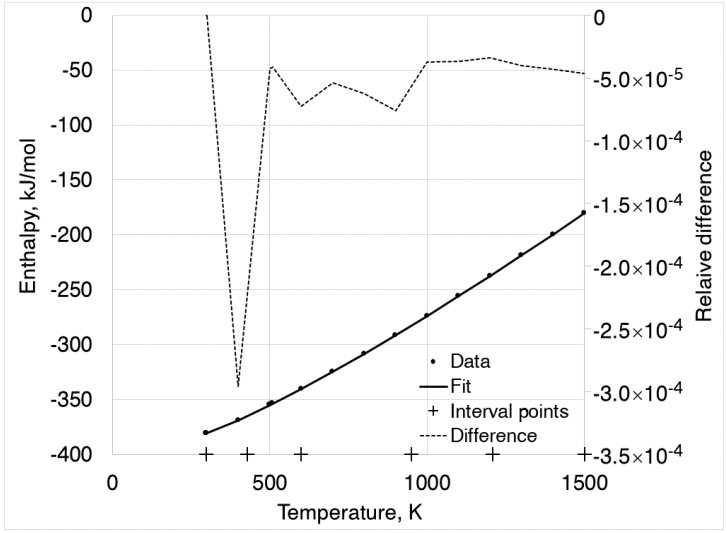


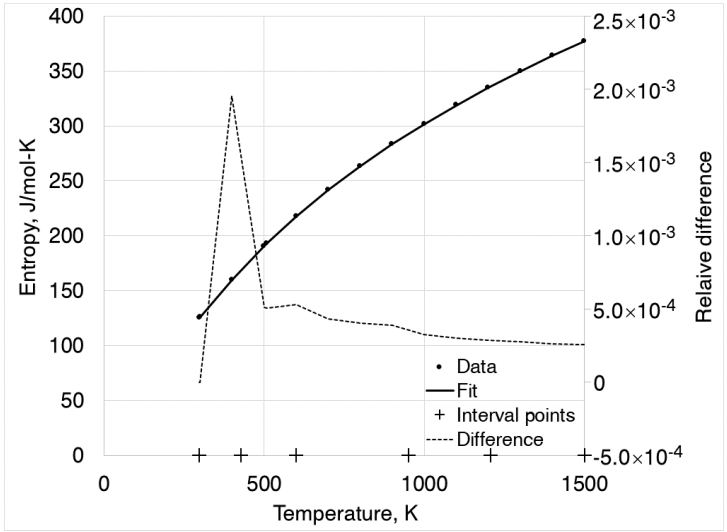


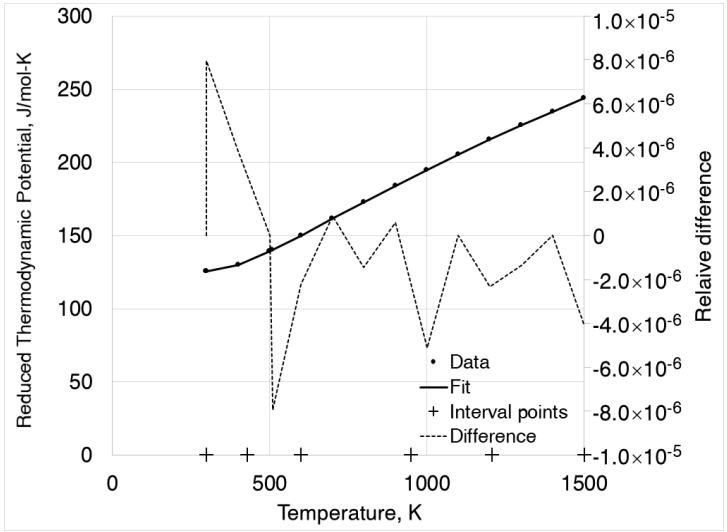
Lithium Perchlorate ($LiClO_4$) – Cl1Li1O4(cr)

MW = 106.392 g/mol – $\Delta_f H^o$(298.15 K) = -380.744 kJ/mol – S^o(298.15 K) = 125.52 J/mol-K – p^o = 0.1 MPa



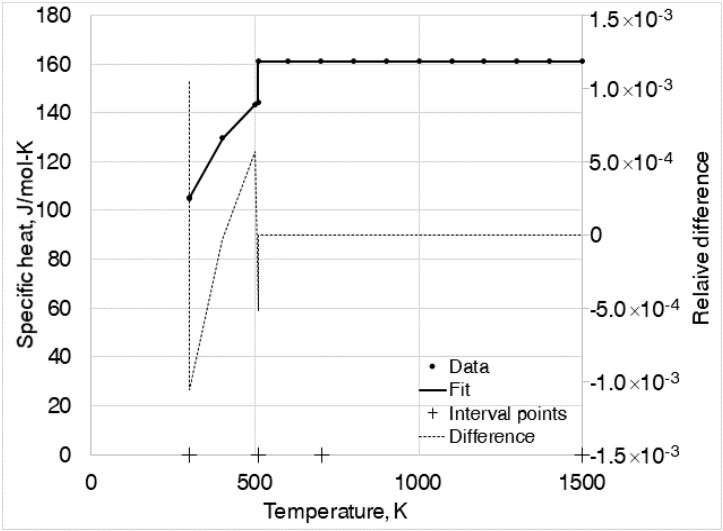


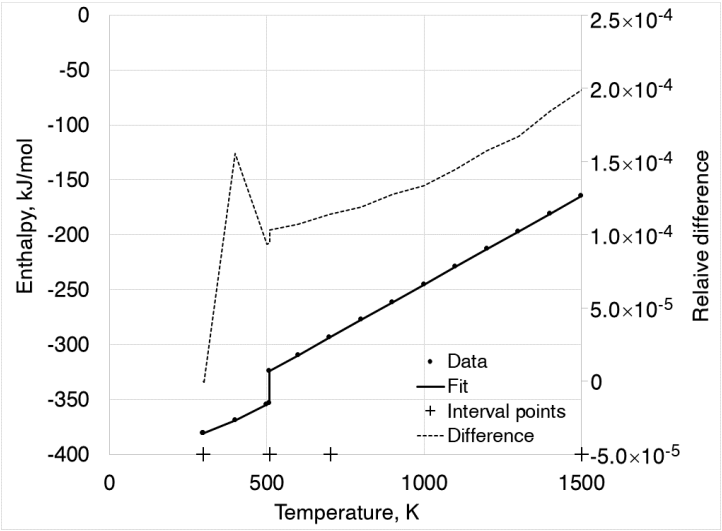


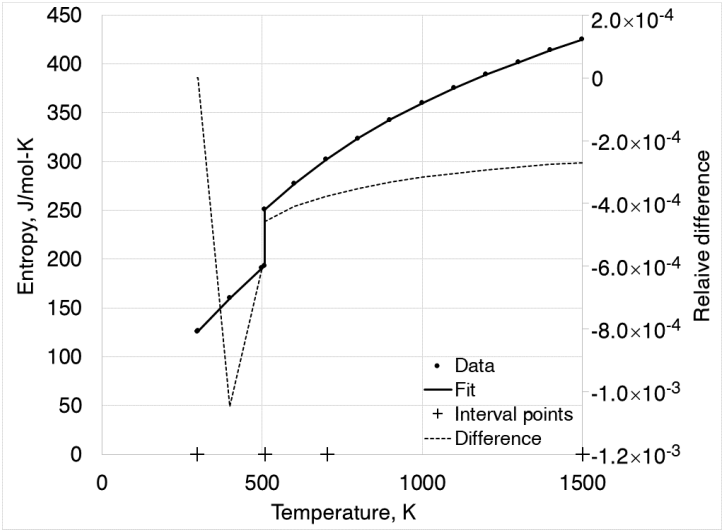


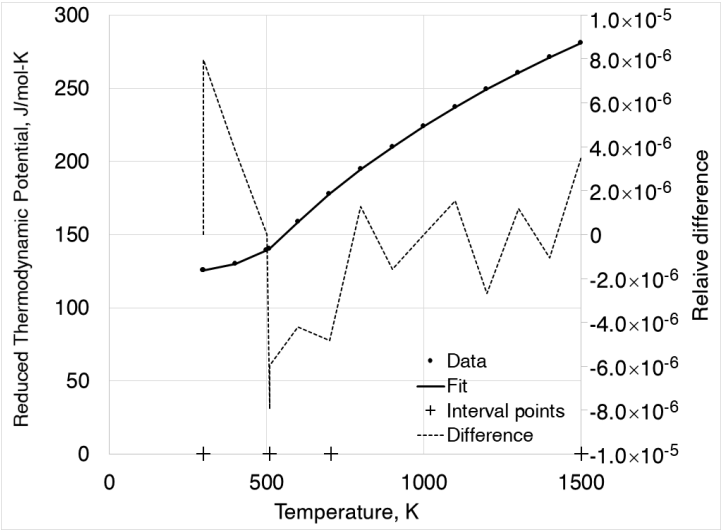
Lithium Perchlorate ($LiClO_4$) – Cl1Li1O4(cr,l)

MW = 106.392 g/mol – $\Delta_f H^o$(298.15 K) = -380.744 kJ/mol – S^o(298.15 K) = 125.52 J/mol-K – p^o = 0.1 MPa



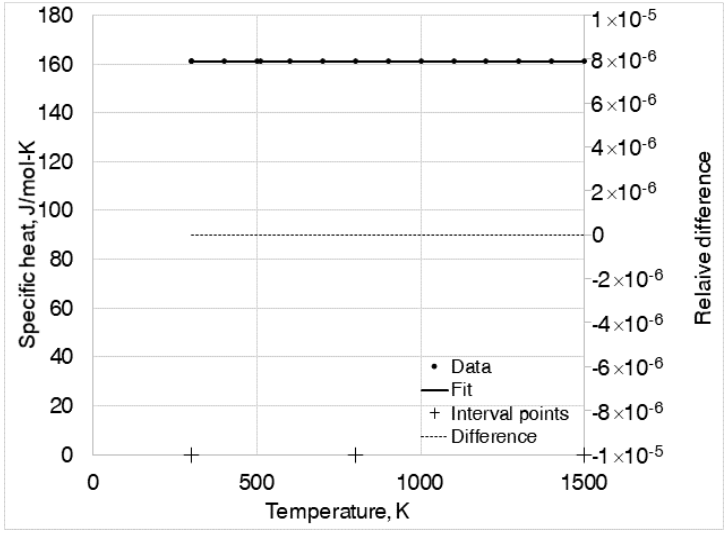


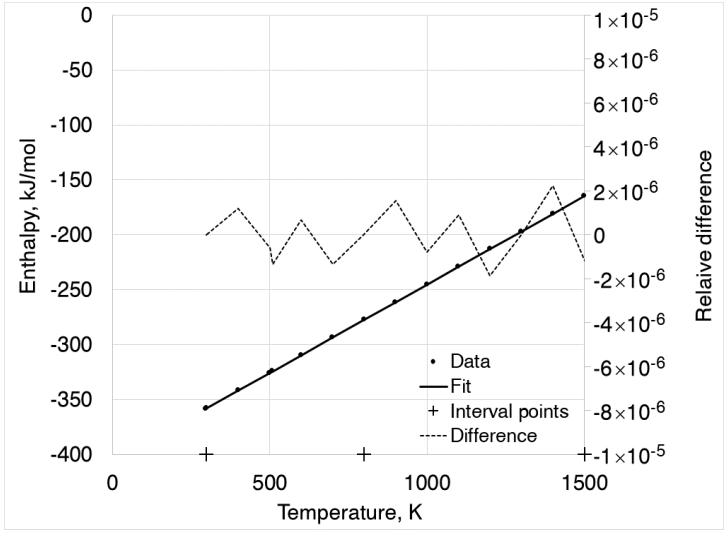


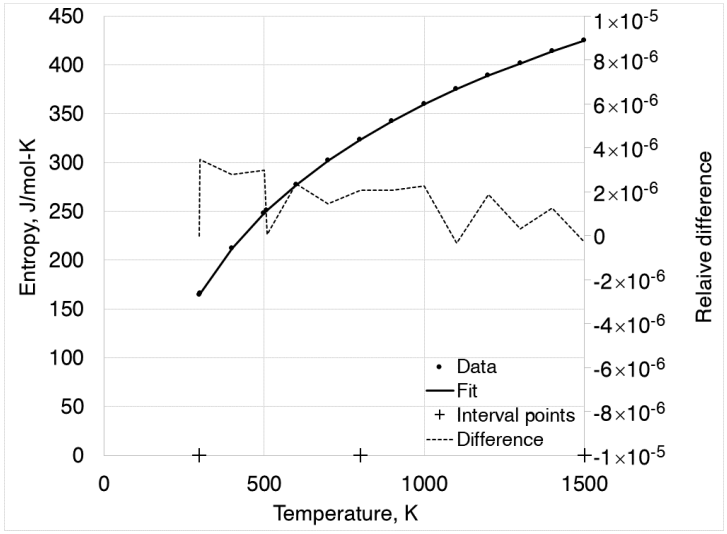


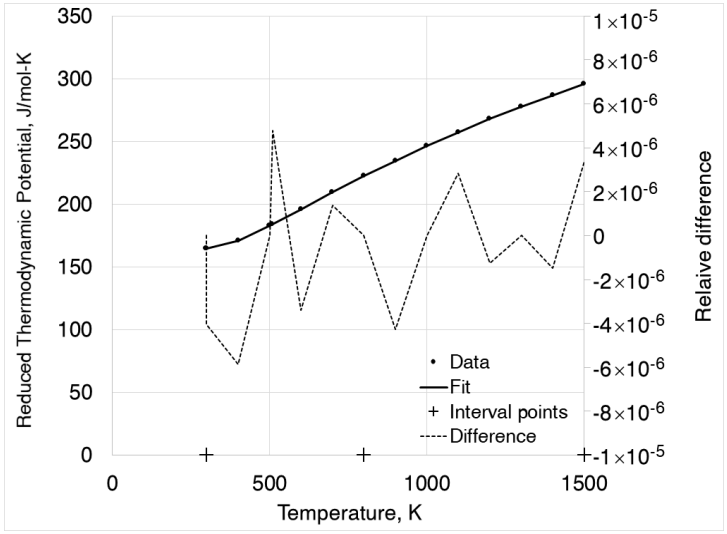
Lithium Perchlorate ($LiClO_4$) – Cl1Li1O4(l)

MW = 106.392 g/mol – $\Delta_f H^o$(298.15 K) = -358.355 kJ/mol – S^o(298.15 K) = 164.611 J/mol-K – p^o = 0.1 MPa



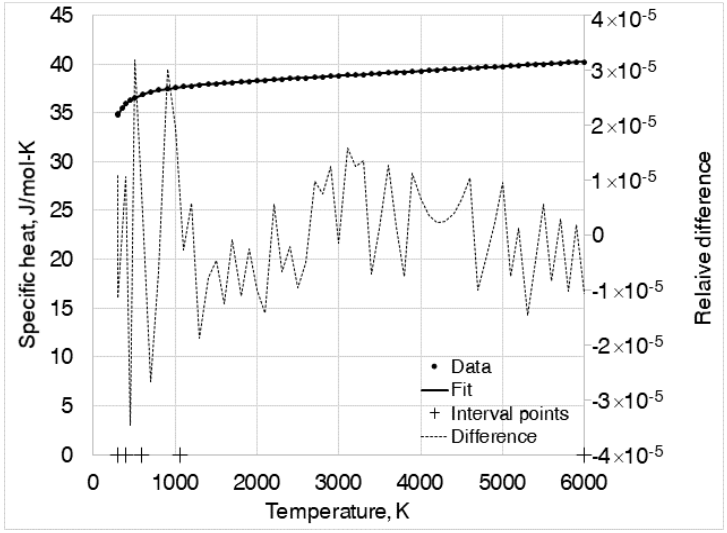


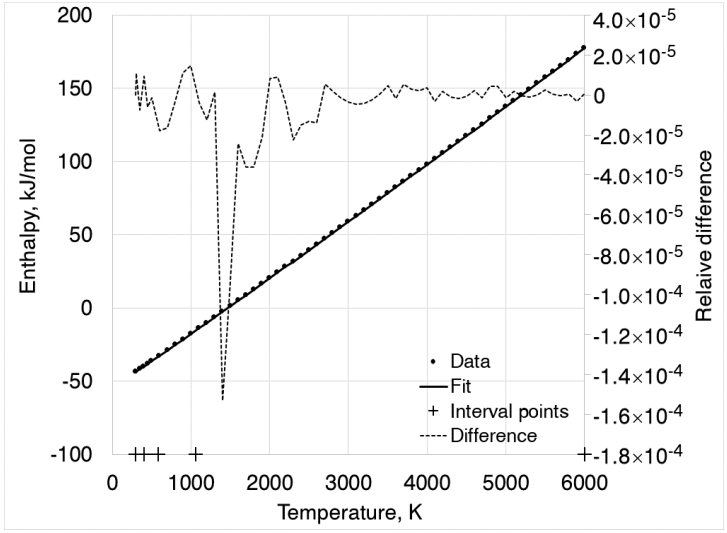


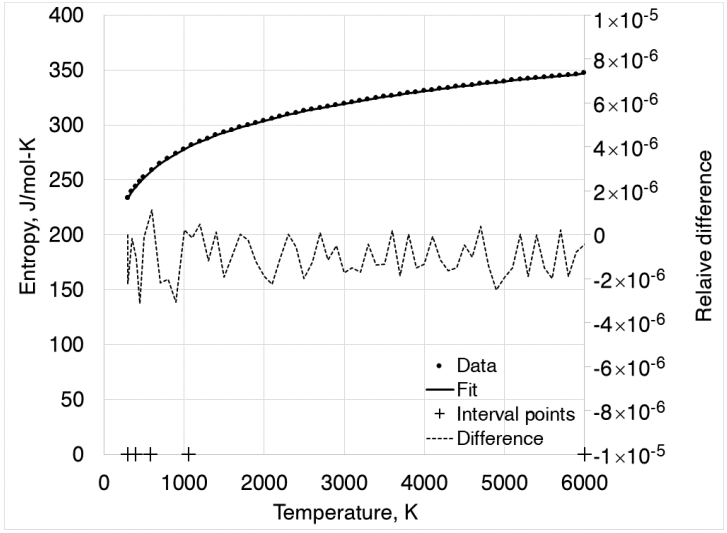


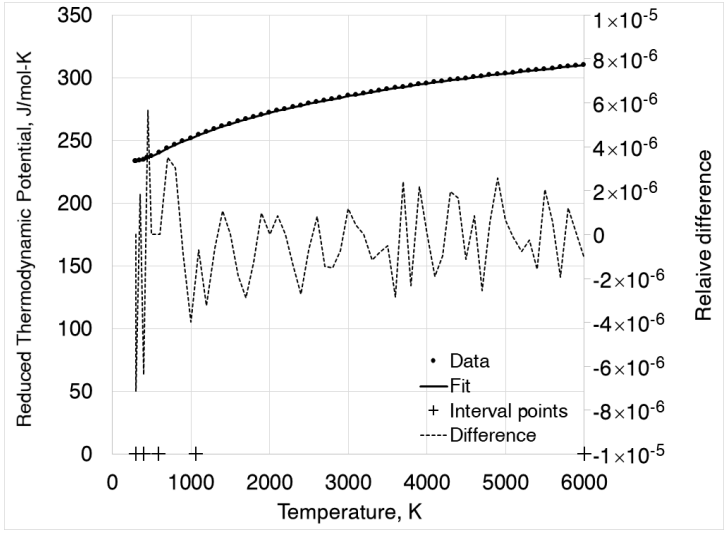
Magnesium Chloride (MgCl) – Cl1Mg1(g)

MW = 59.758 g/mol – $\Delta_f H^o$(298.15 K) = -43.514 kJ/mol – S^o(298.15 K) = 233.413 J/mol-K – p^o = 0.1 MPa



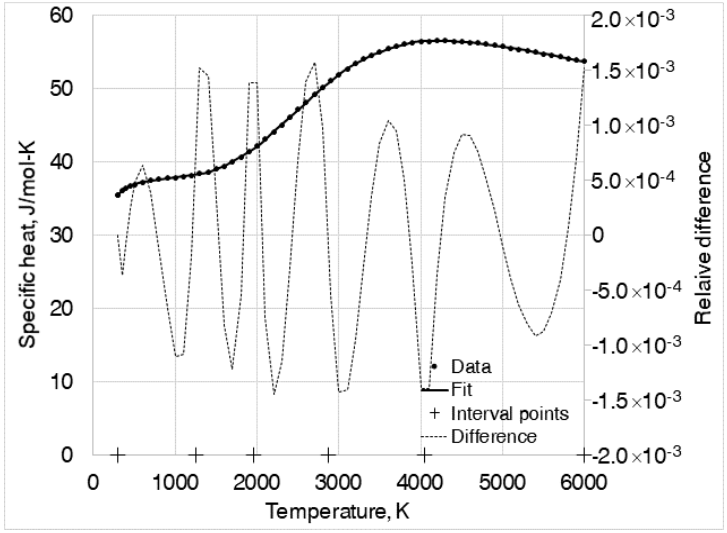


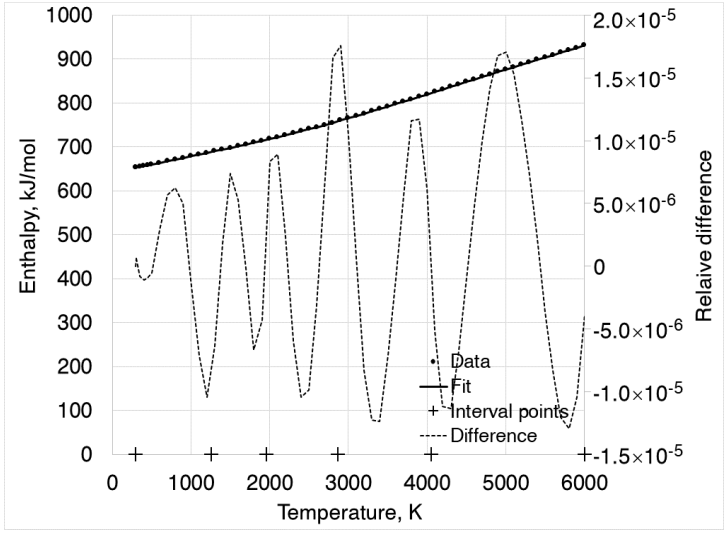


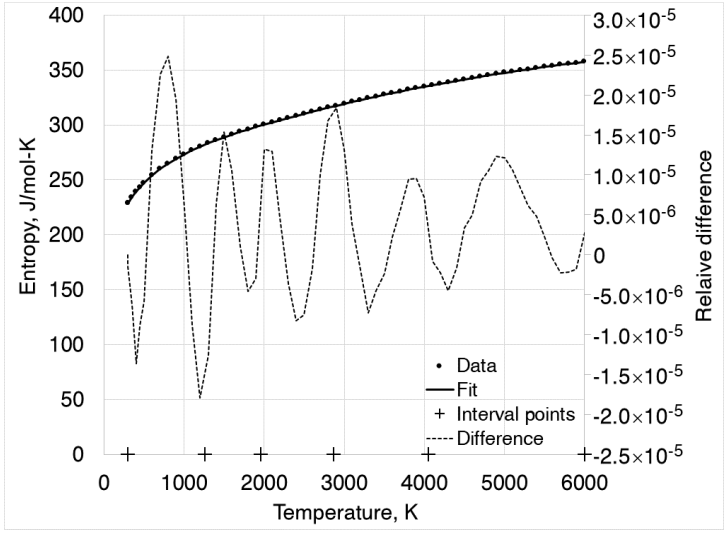


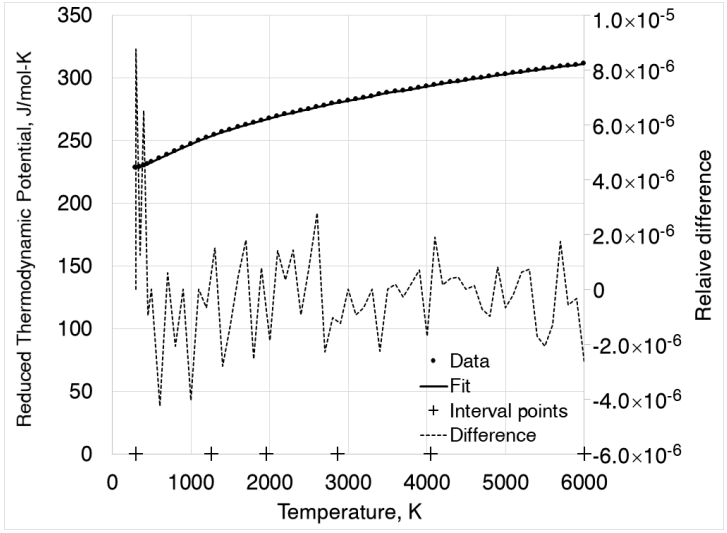
Magnesium Chloride, Ion ($MgCl^+$) – Cl1Mg1+(g)

MW = 59.757 g/mol – $\Delta_f H^o$(298.15 K) = 652.704 kJ/mol – S^o(298.15 K) = 228.556 J/mol-K – p^o = 0.1 MPa



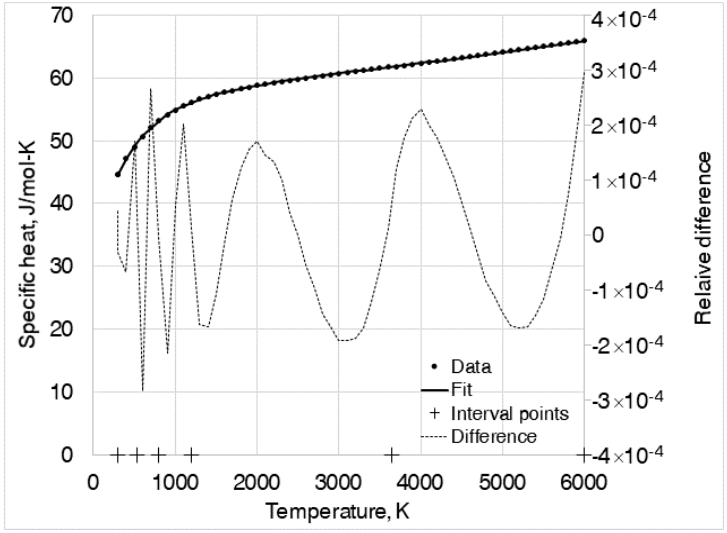


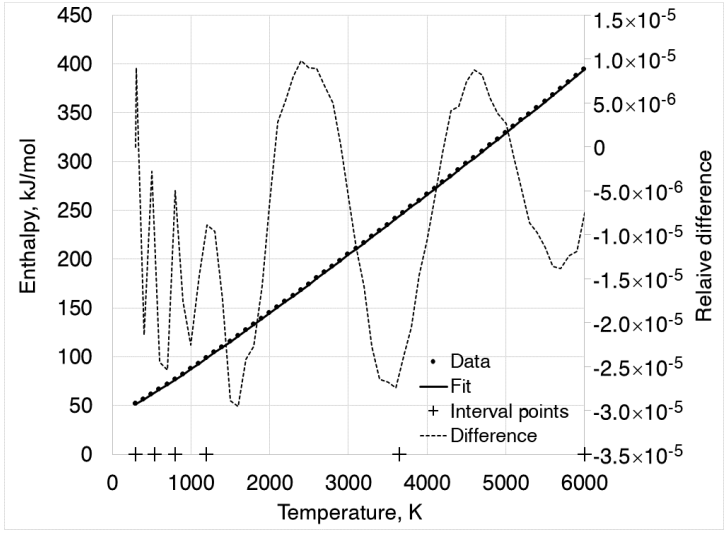


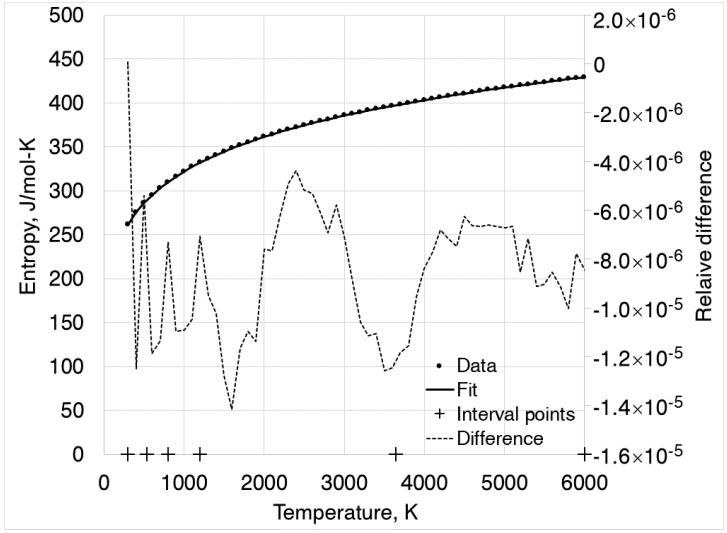


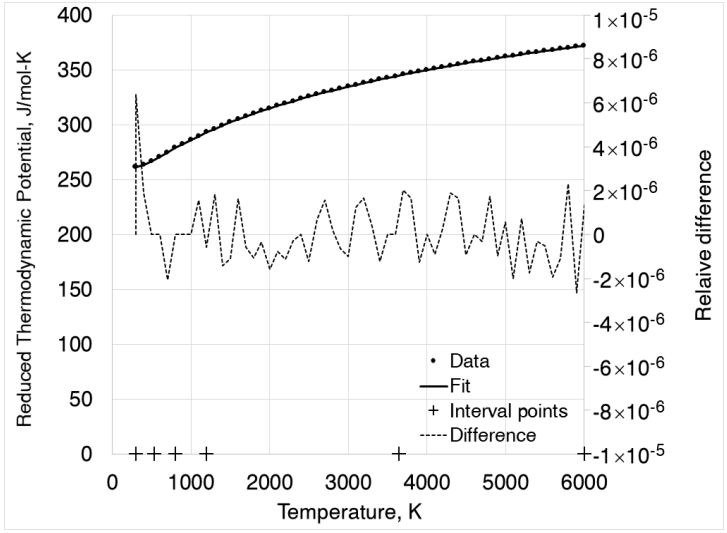
Nitrosyl Chloride (ONCl) – Cl1N1O1(g)

MW = 65.459 g/mol – $\Delta_f H^o$(298.15 K) = 51.714 kJ/mol – S^o(298.15 K) = 261.68 J/mol-K – p^o = 0.1 MPa



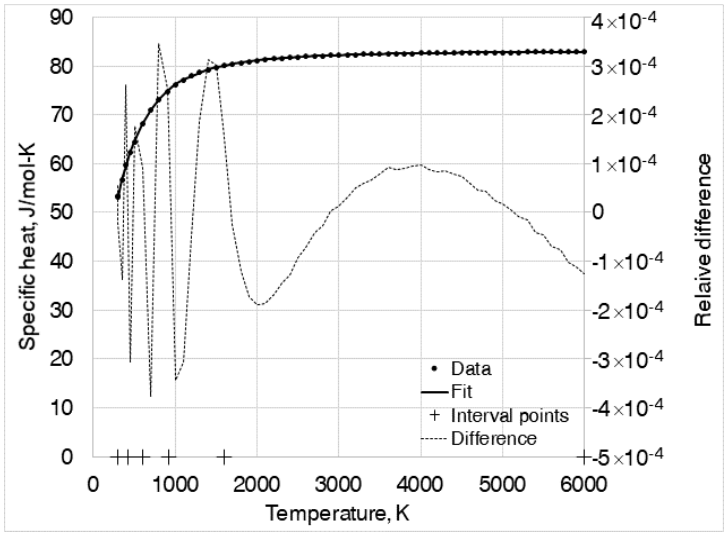


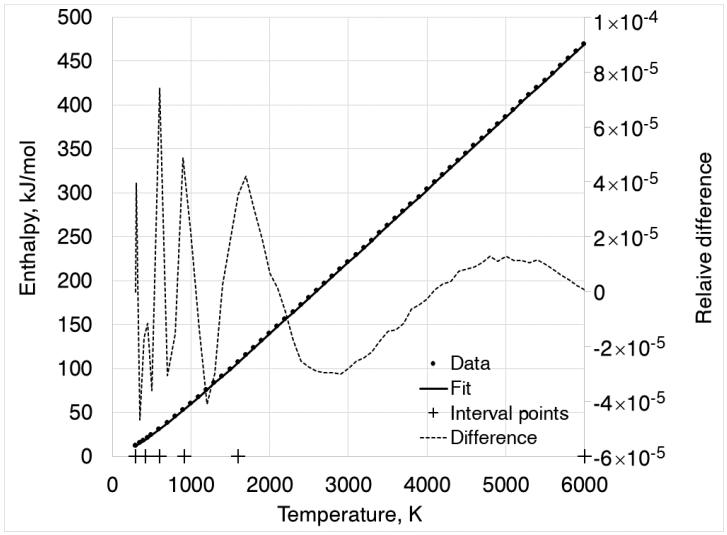


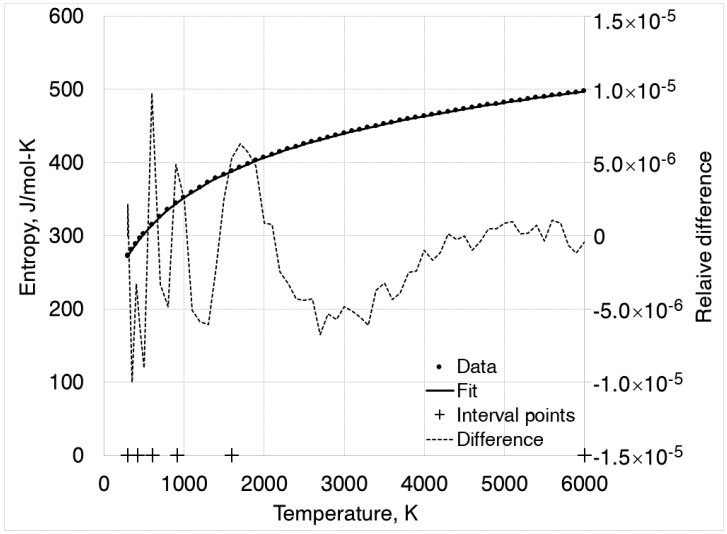


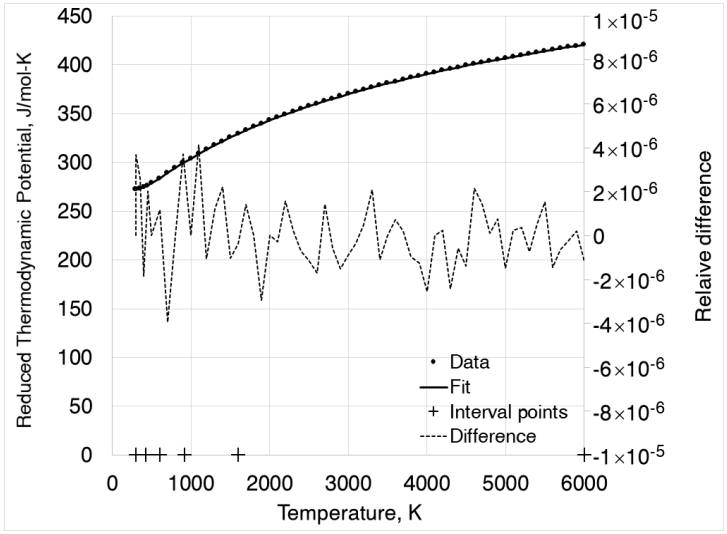
Nitryl Chloride (NO_2Cl) – Cl1N1O2(g)

MW = 81.458 g/mol – $\Delta_f H^o$(298.15 K) = 12.134 kJ/mol – S^o(298.15 K) = 272.187 J/mol-K – p^o = 0.1 MPa



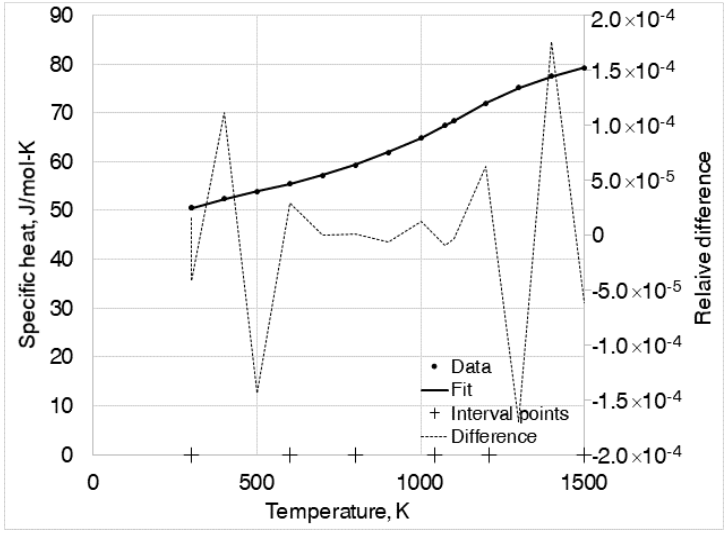


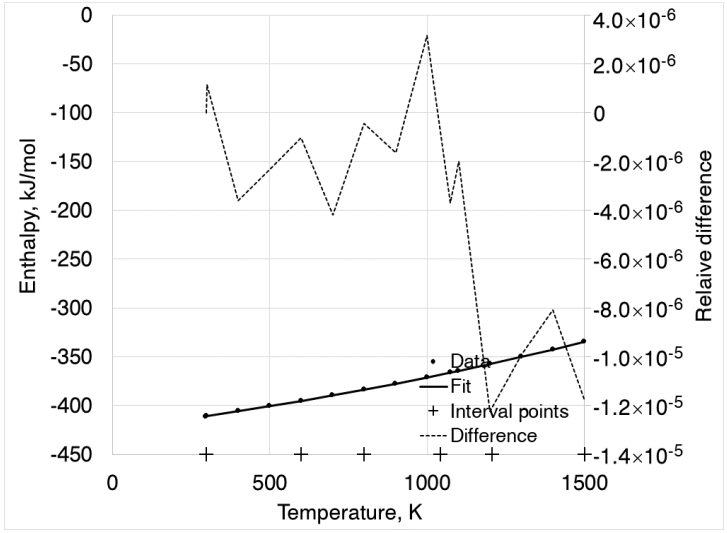


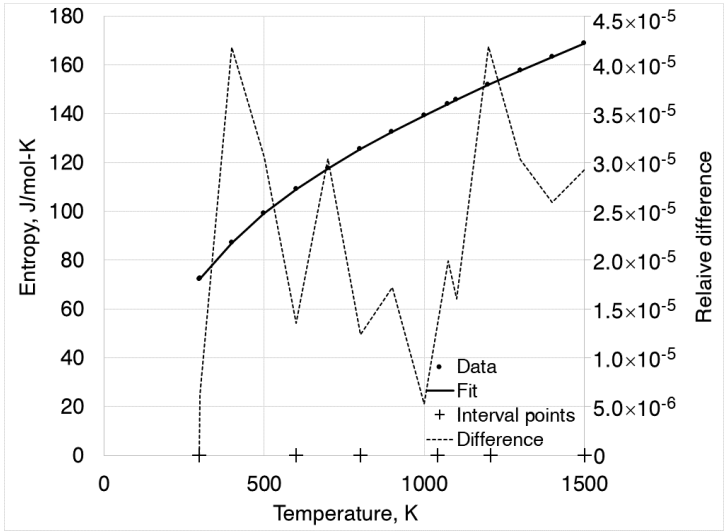


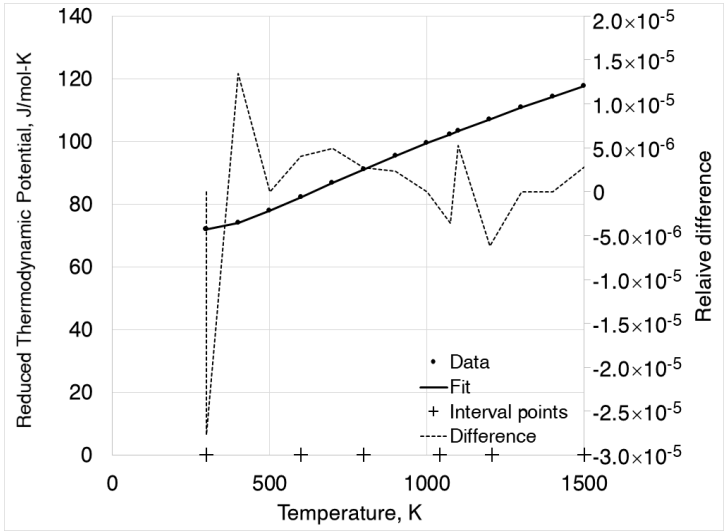
Sodium Chloride (NaCl) – Cl1Na1(cr)

MW = 58.443 g/mol – $\Delta_f H^o$(298.15 K) = -411.12 kJ/mol – S^o(298.15 K) = 72.115 J/mol-K – p^o = 0.1 MPa



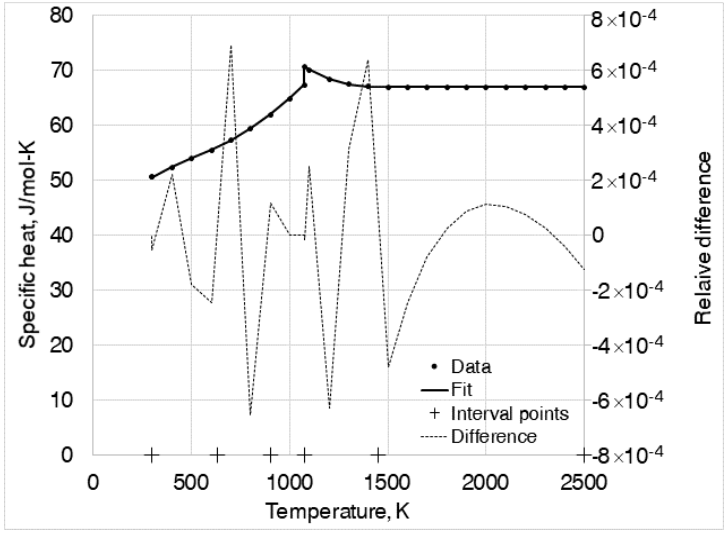


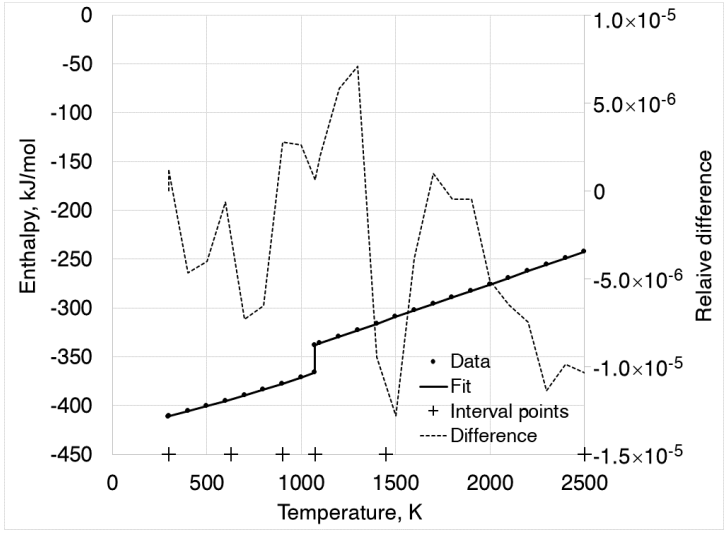


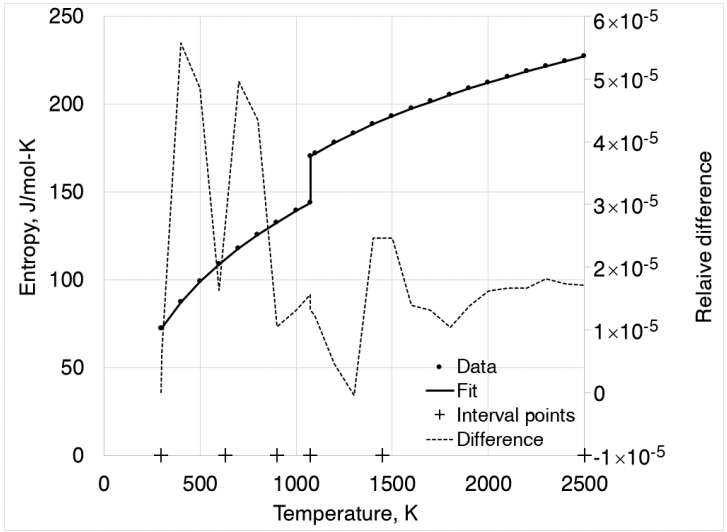


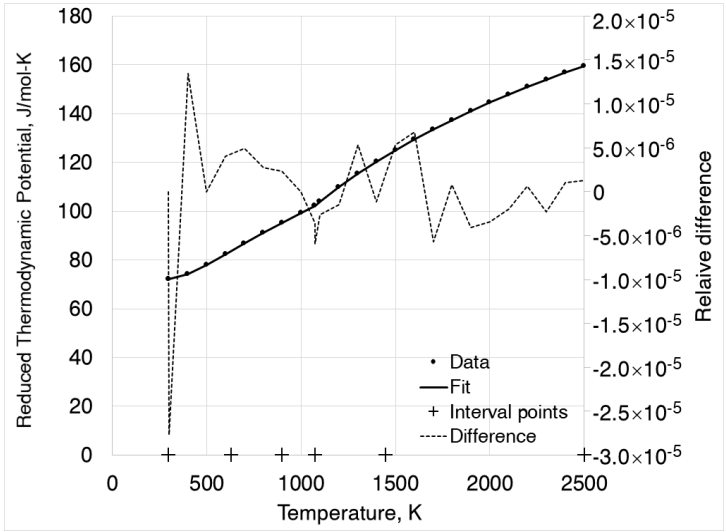
Sodium Chloride (NaCl) – Cl1Na1(cr,l)

MW = 58.443 g/mol – $\Delta_f H^o$(298.15 K) = -411.12 kJ/mol – S^o(298.15 K) = 72.115 J/mol-K – p^o = 0.1 MPa



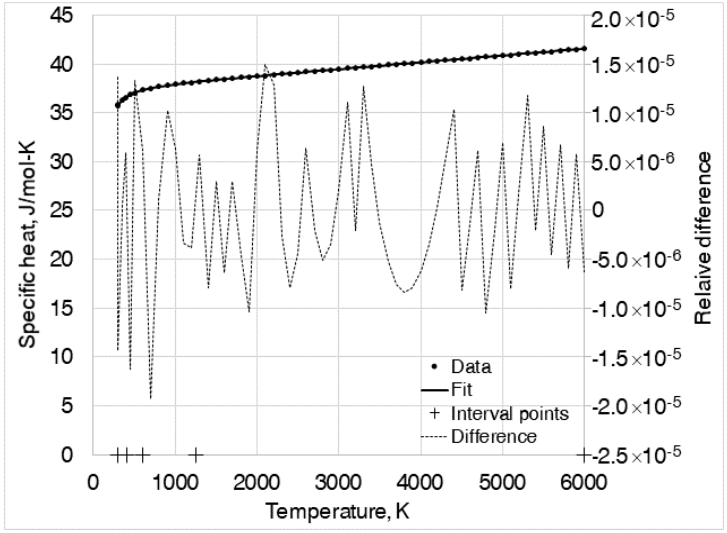


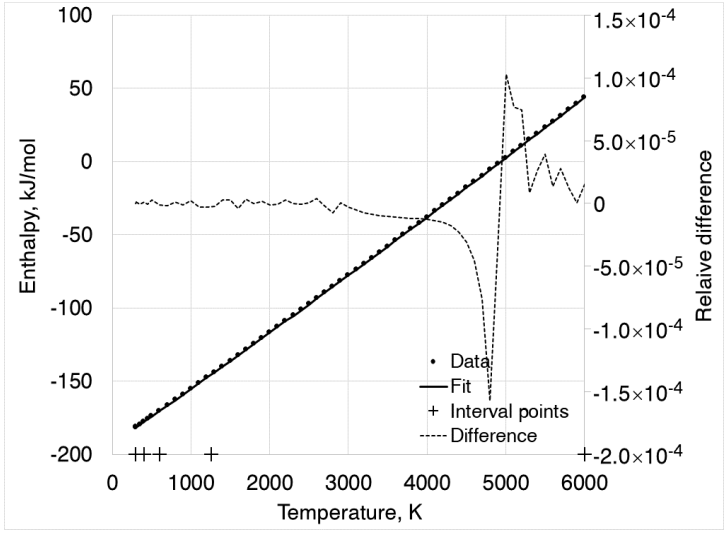


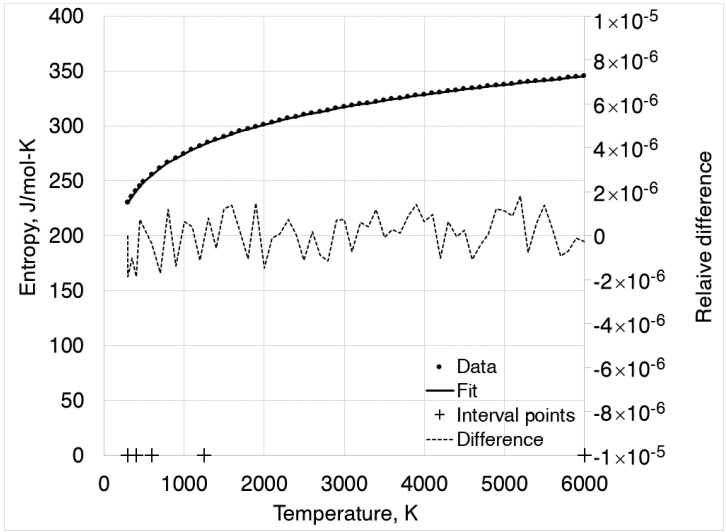


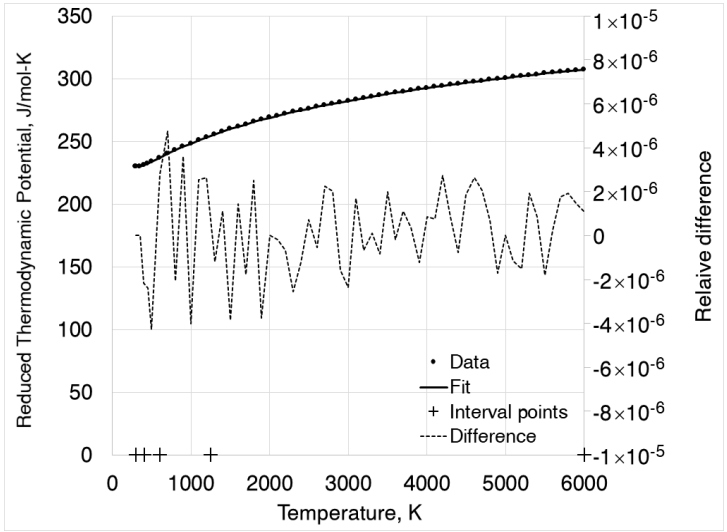
Sodium Chloride (NaCl) – Cl1Na1(g)

MW = 58.443 g/mol – $\Delta_f H^o$(298.15 K) = -181.418 kJ/mol – S^o(298.15 K) = 229.793 J/mol-K – p^o = 0.1 MPa



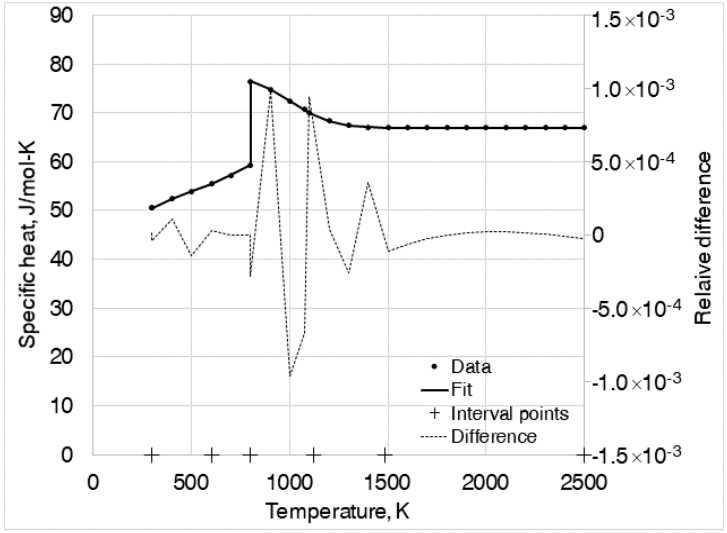


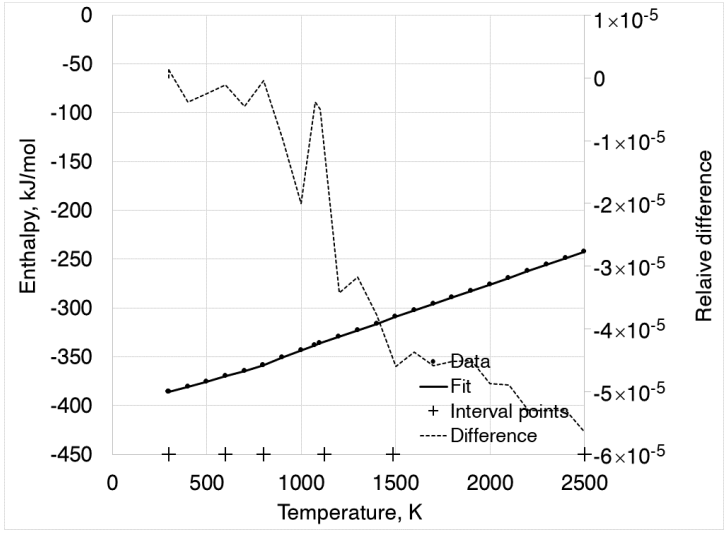


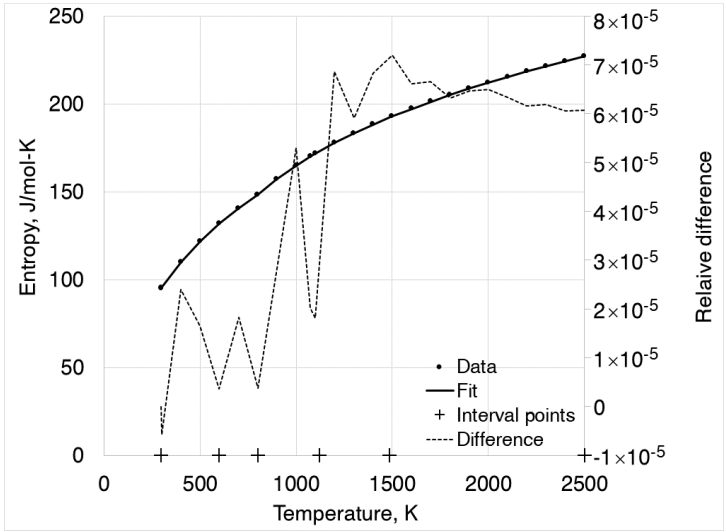


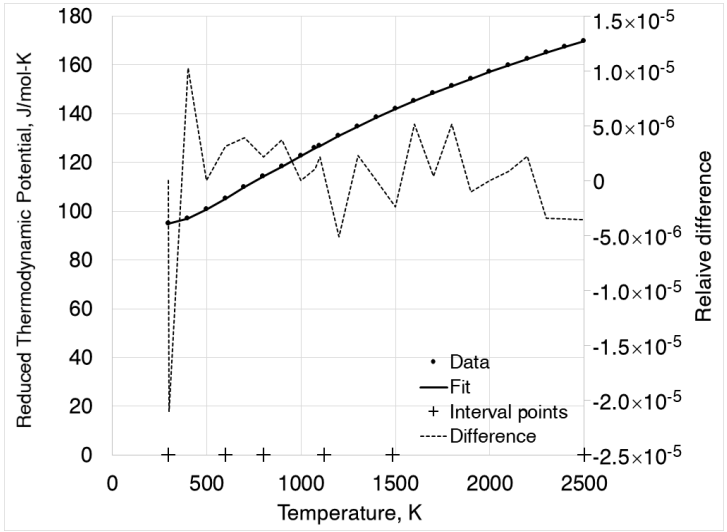
Sodium Chloride (NaCl) – Cl1Na1(l)

MW = 58.443 g/mol – $\Delta_f H^o$(298.15 K) = -385.923 kJ/mol – S^o(298.15 K) = 95.055 J/mol-K – p^o = 0.1 MPa



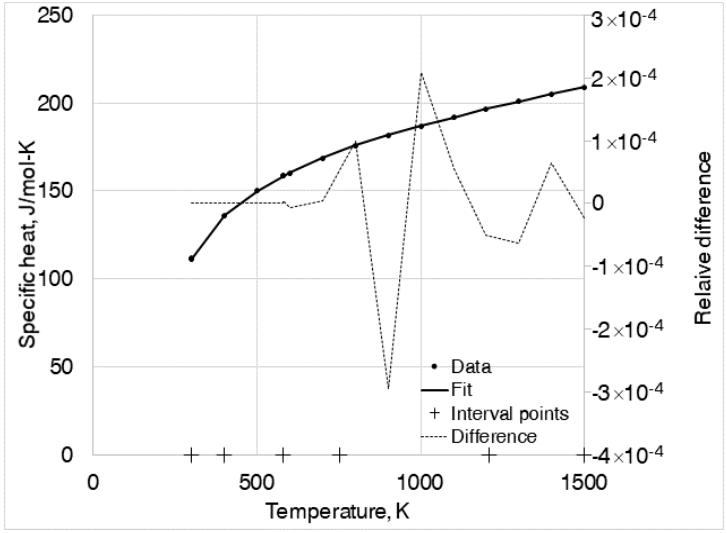


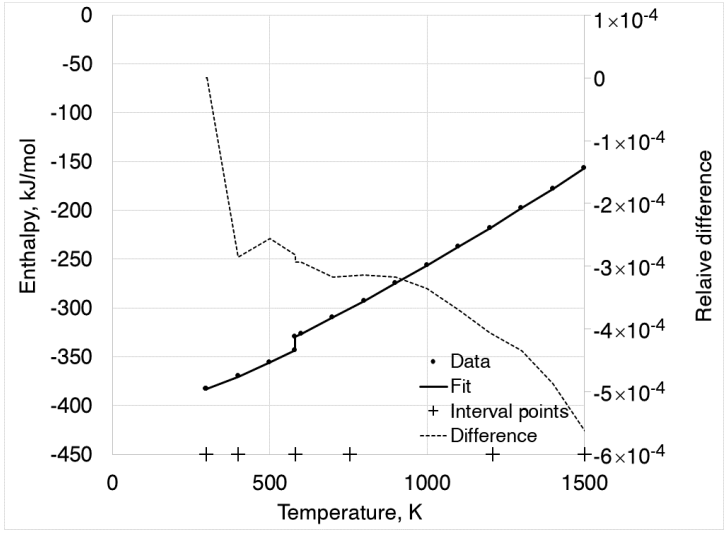


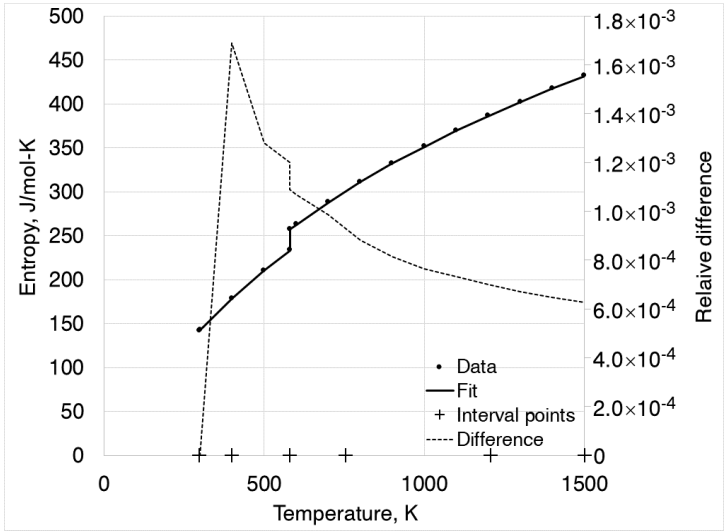


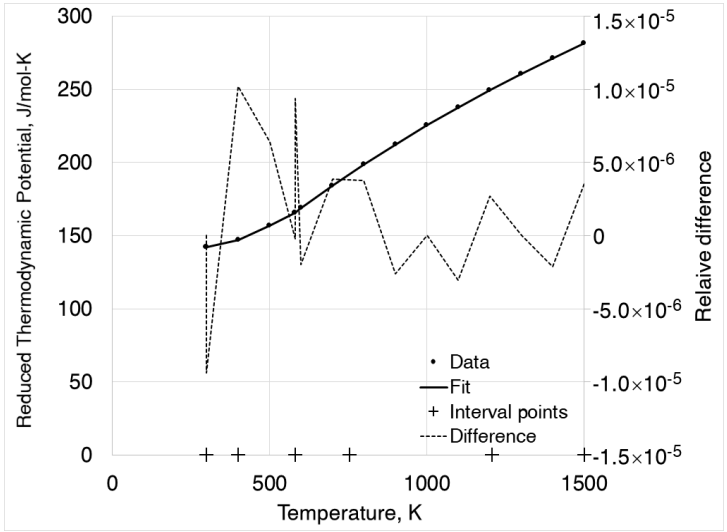
Sodium Perchlorate ($NaClO_4$) – Cl1Na1O4(cr)

MW = 122.44 g/mol – $\Delta_f H^o$(298.15 K) = -382.752 kJ/mol – S^o(298.15 K) = 142.256 J/mol-K – p^o = 0.1 MPa



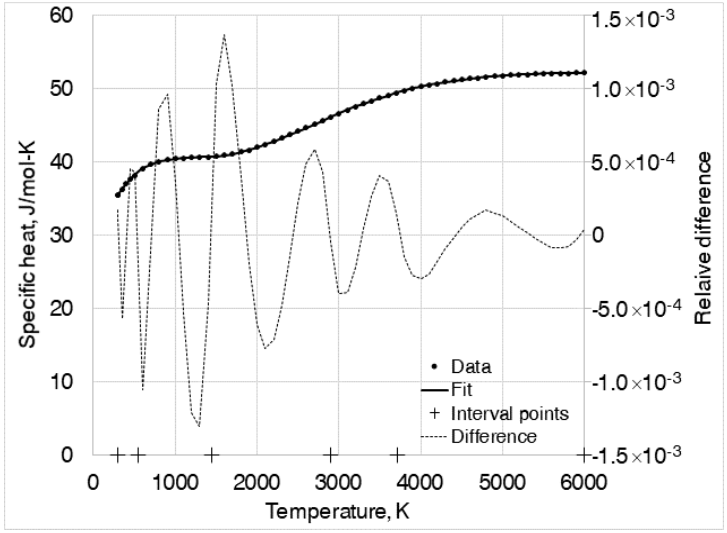


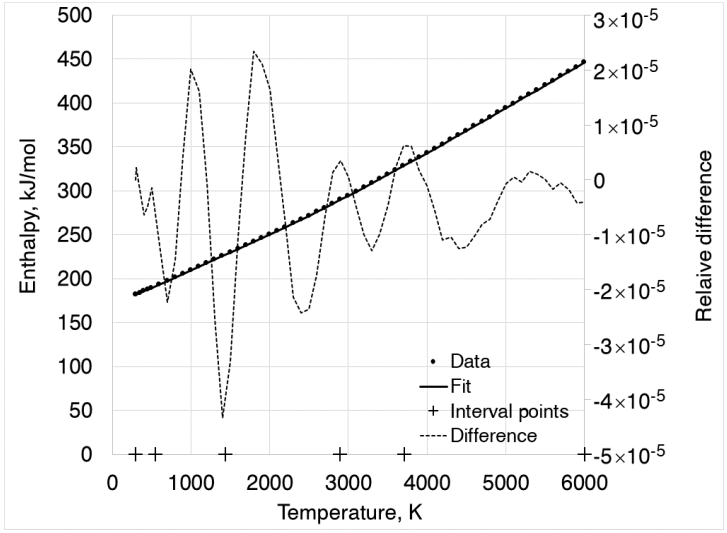


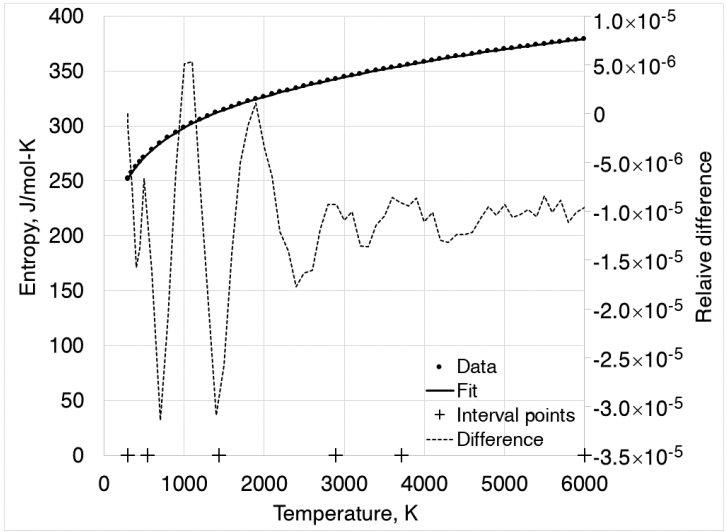


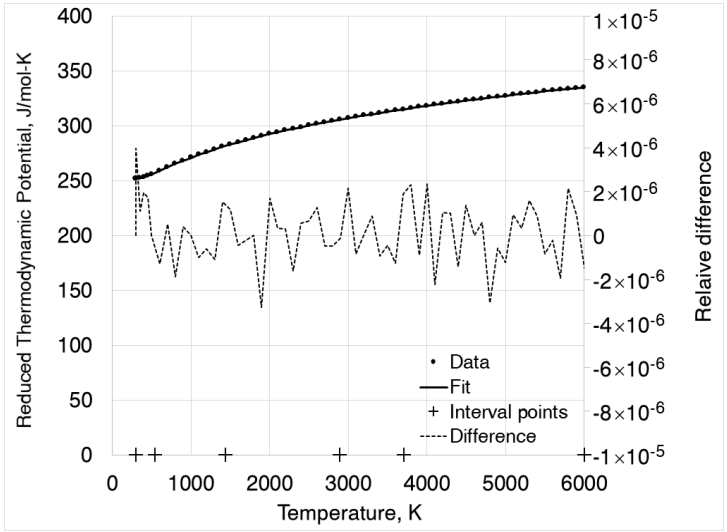
Nickel Chloride (NiCl) – Cl1Ni1(g)

MW = 94.14 g/mol – $\Delta_f H^o$(298.15 K) = 182.004 kJ/mol – S^o(298.15 K) = 251.883 J/mol-K – p^o = 0.1 MPa



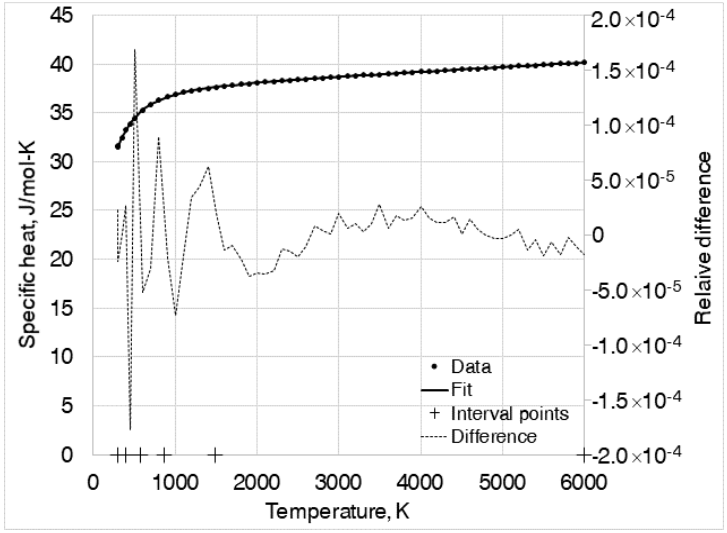


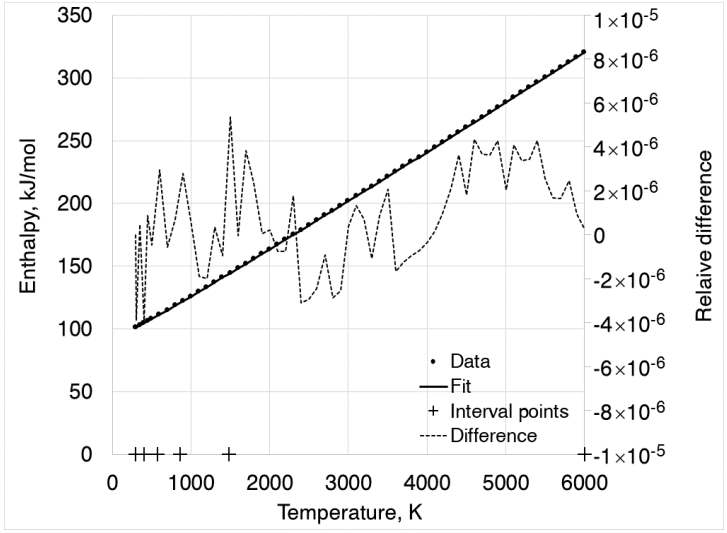


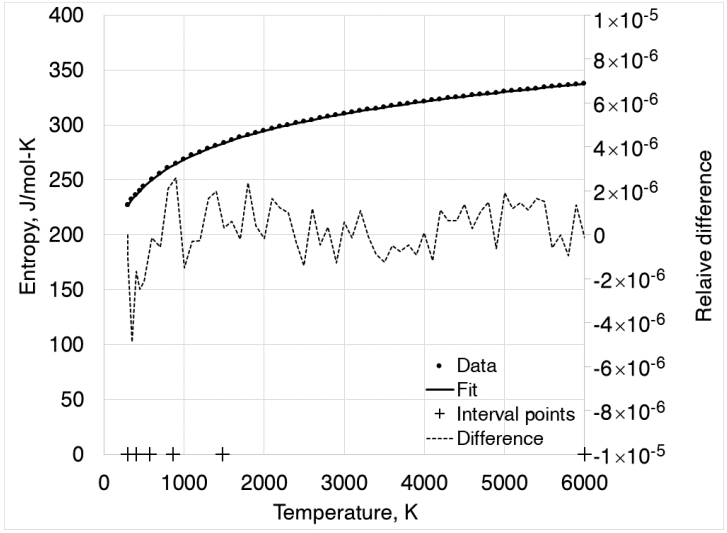


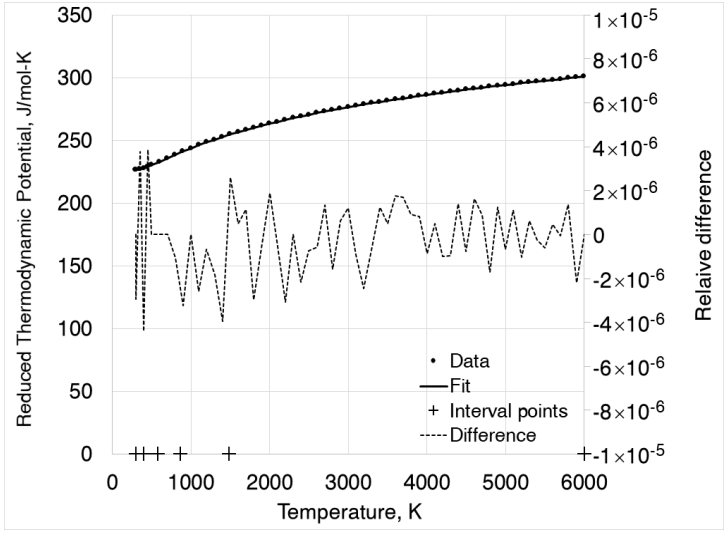
Chlorine Oxide (ClO) – Cl1O1(g)

MW = 51.452 g/mol – $\Delta_f H^o$(298.15 K) = 101.219 kJ/mol – S^o(298.15 K) = 226.648 J/mol-K – p^o = 0.1 MPa



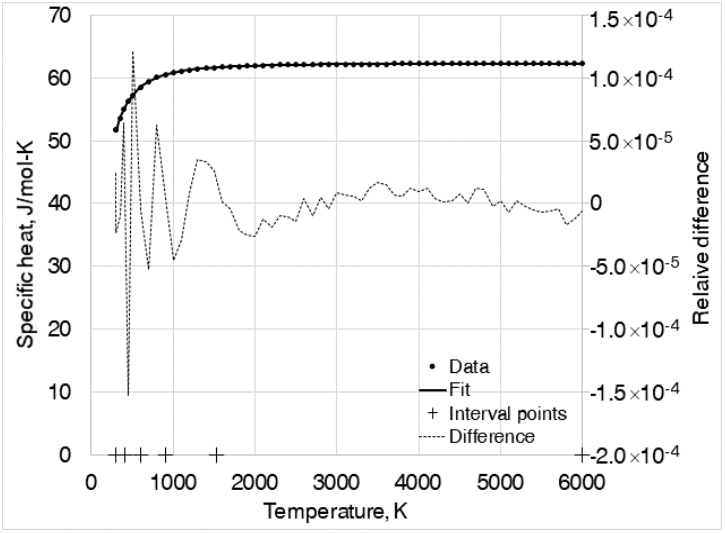


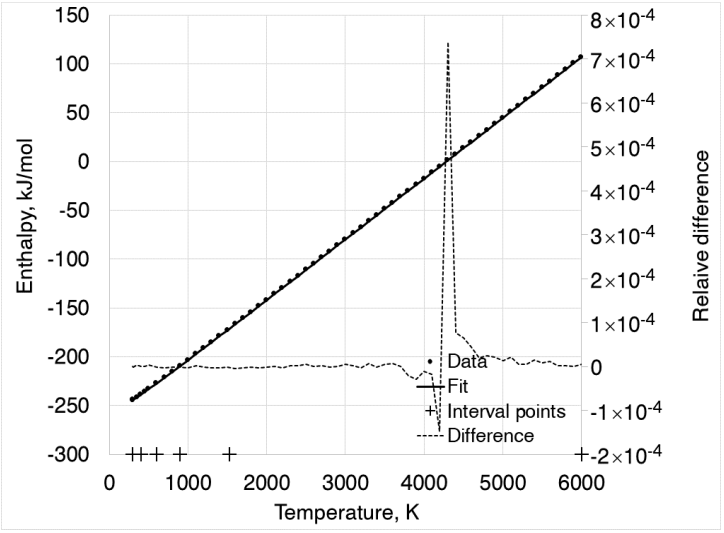


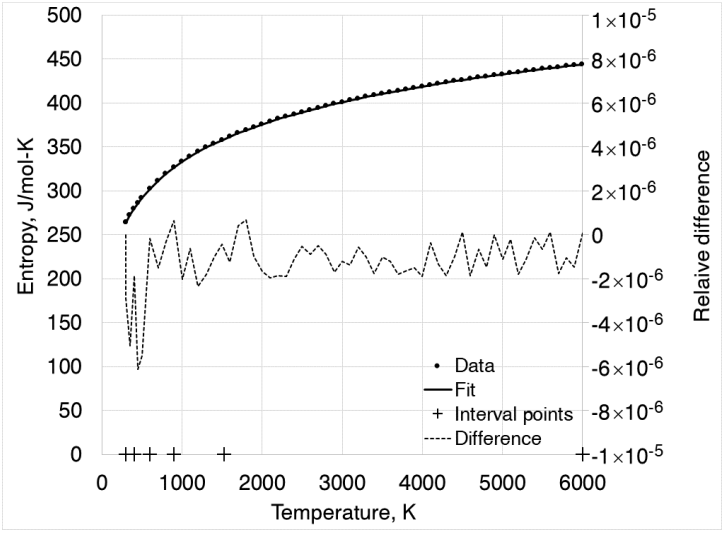


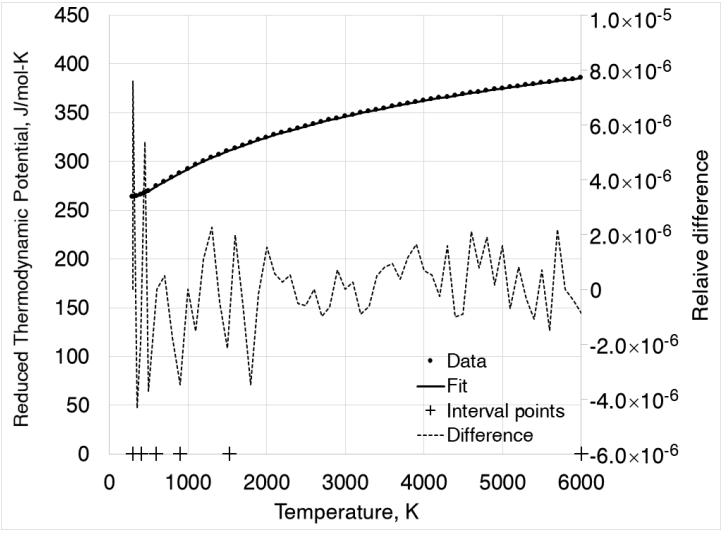
Titanium Chloride Oxide (OTiCl) – Cl1O1Ti1(g)

MW = 99.33 g/mol – $\Delta_f H^o$(298.15 K) = -244.262 kJ/mol – S^o(298.15 K) = 263.661 J/mol-K – p^o = 0.1 MPa



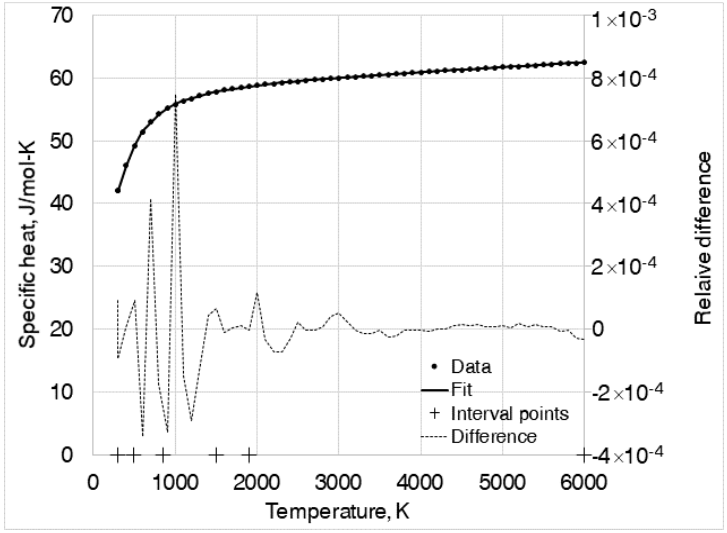


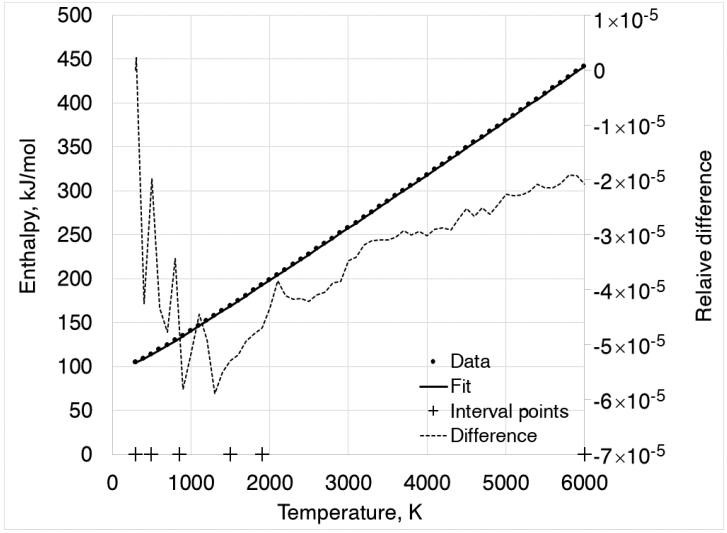


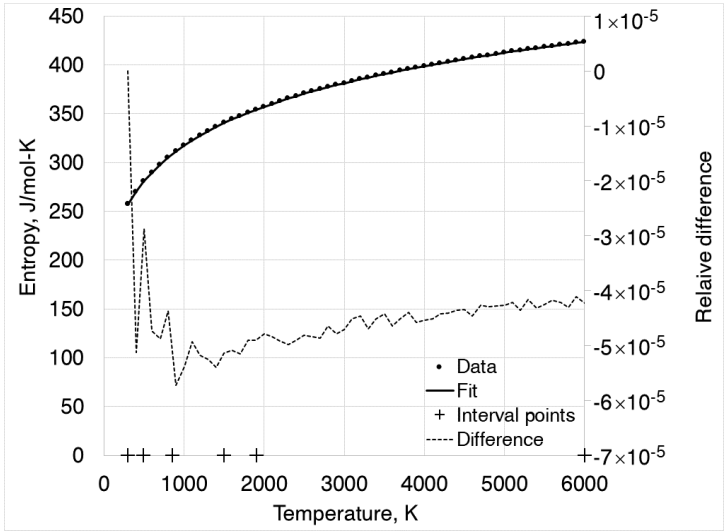


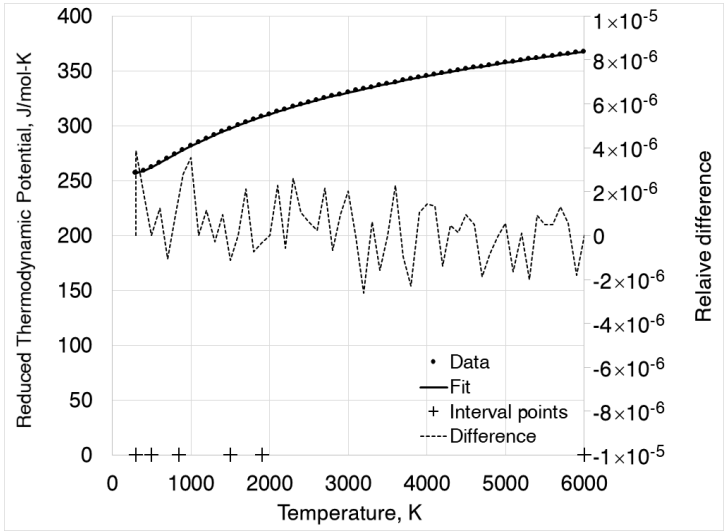
Chlorine Oxide (ClO_2) – Cl1O2(g)

MW = 67.452 g/mol – $\Delta_f H^o$(298.15 K) = 104.6 kJ/mol – S^o(298.15 K) = 257.229 J/mol-K – p^o = 0.1 MPa



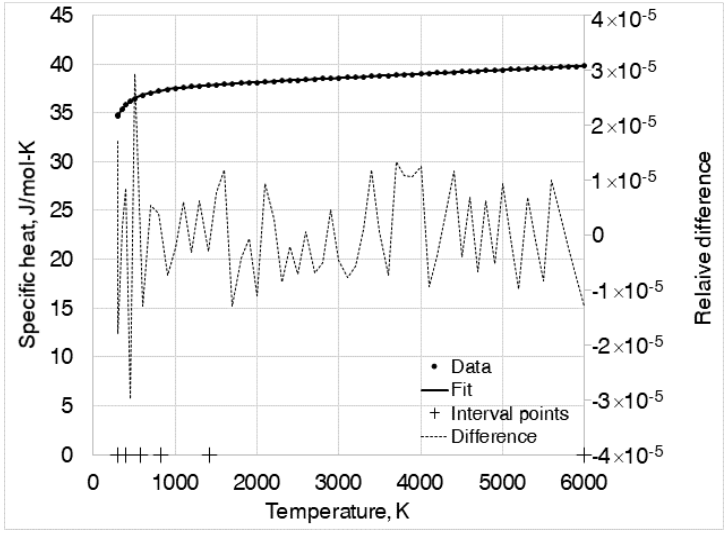


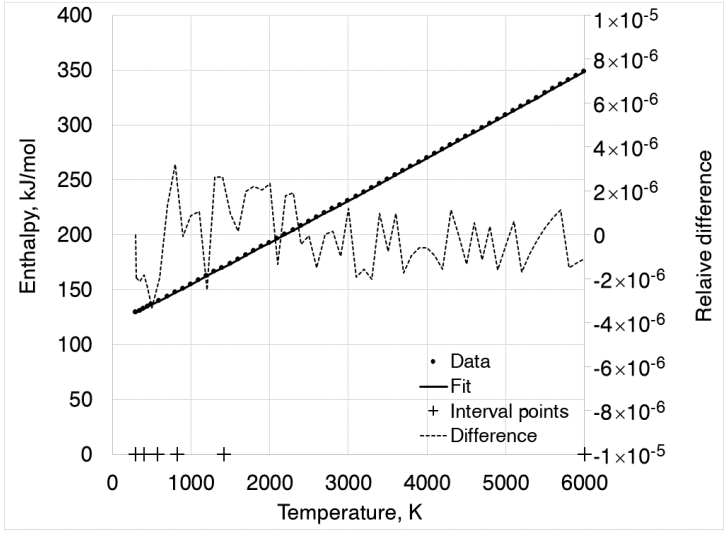


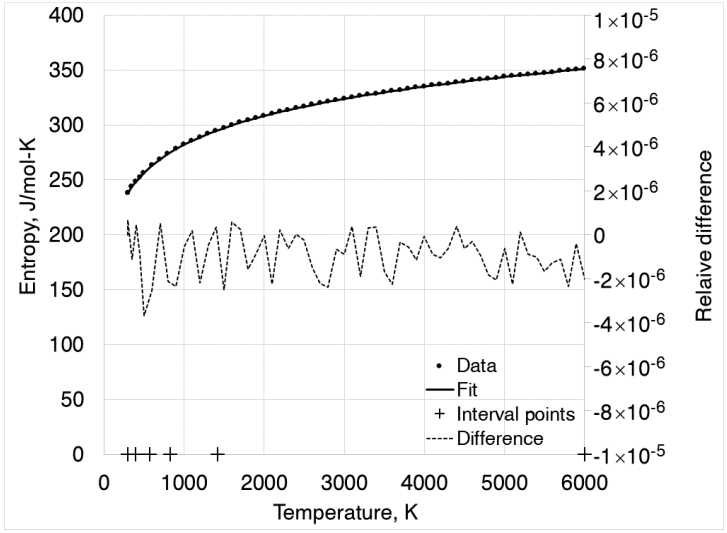


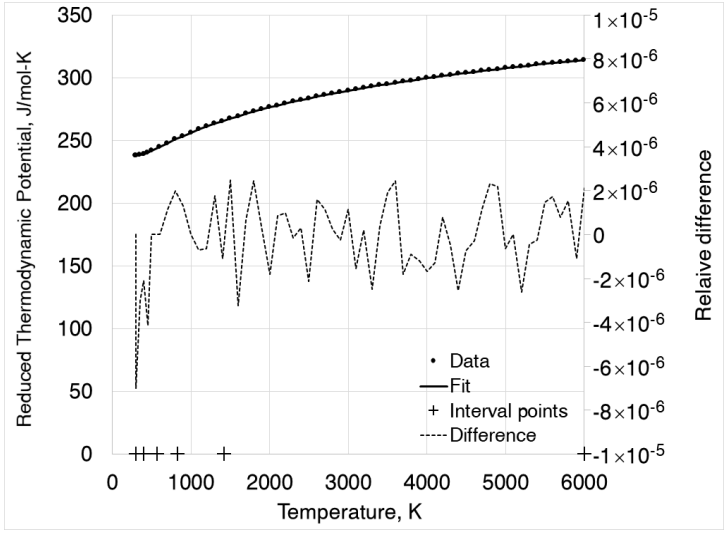
Phosphorus Chloride (PCl) – Cl1P1(g)

MW = 66.427 g/mol – $\Delta_f H^o$(298.15 K) = 129 kJ/mol – S^o(298.15 K) = 237.879 J/mol-K – p^o = 0.1 MPa



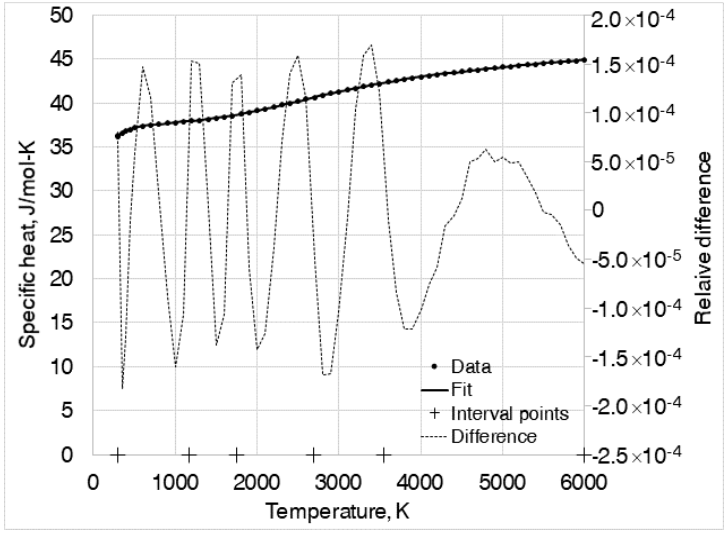


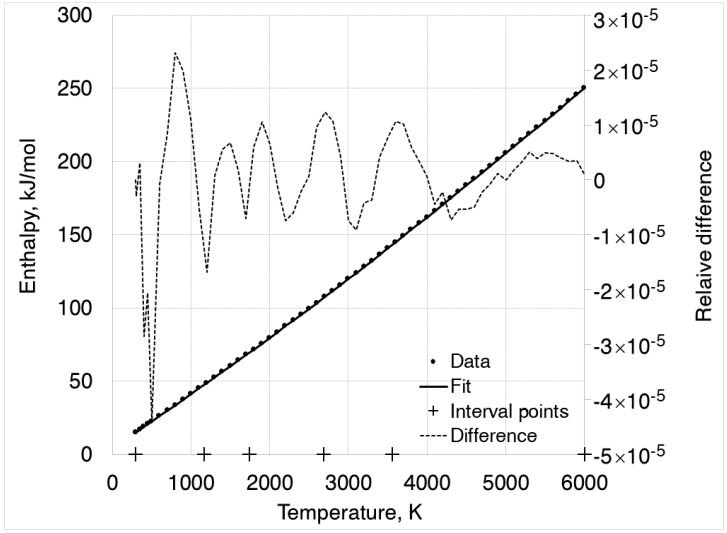


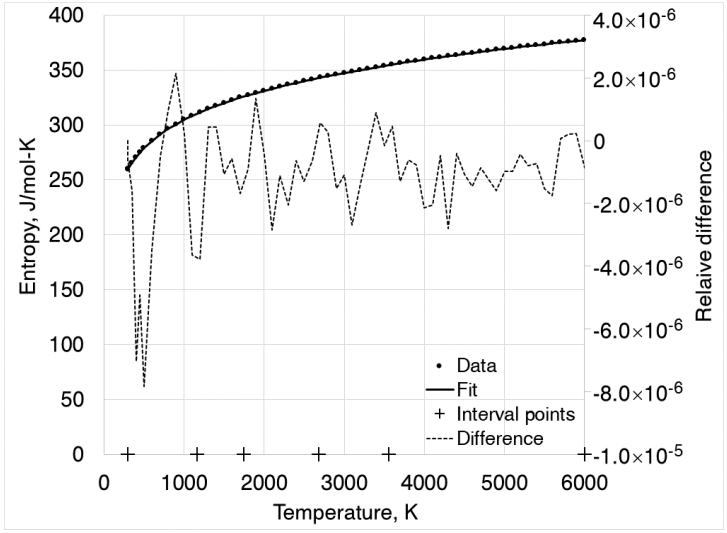


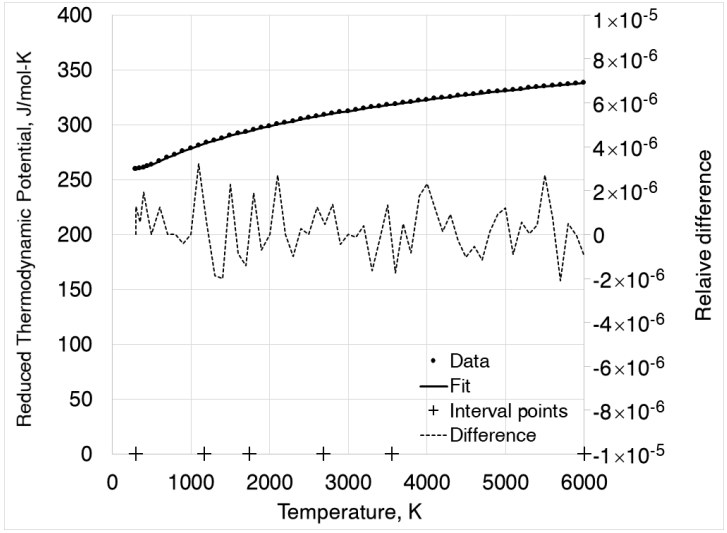
Lead Chloride (PbCl) – Cl1Pb1(g)

MW = 242.7 g/mol – $\Delta_f H^o$(298.15 K) = 15.062 kJ/mol – S^o(298.15 K) = 259.613 J/mol-K – p^o = 0.1 MPa



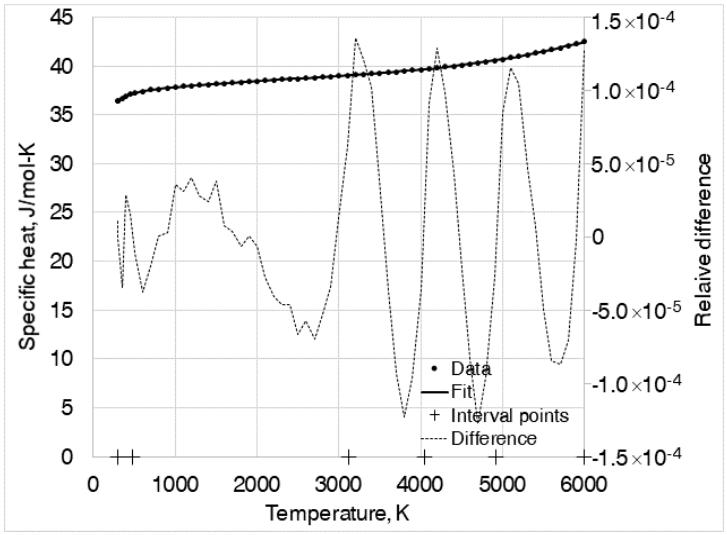


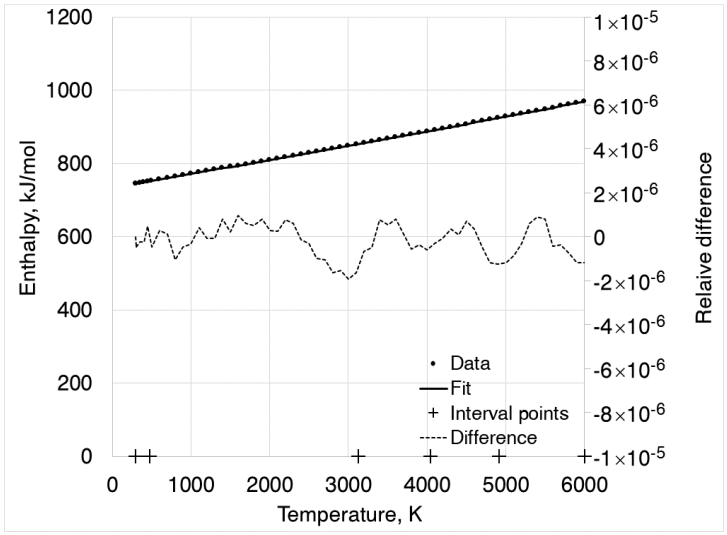


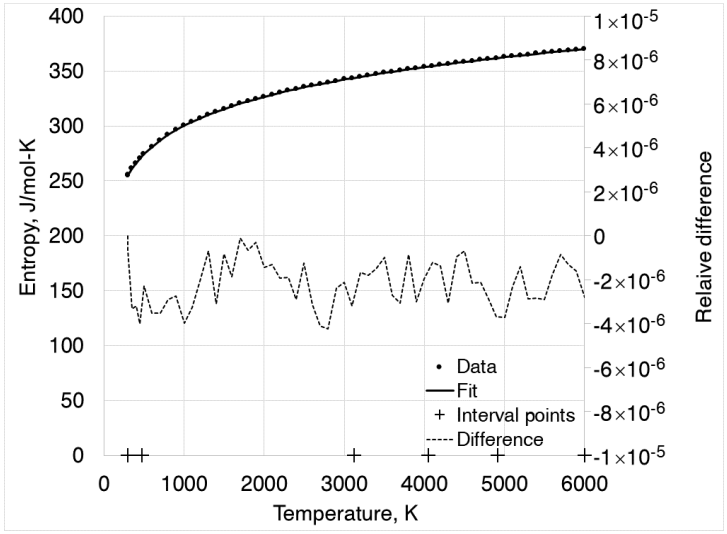


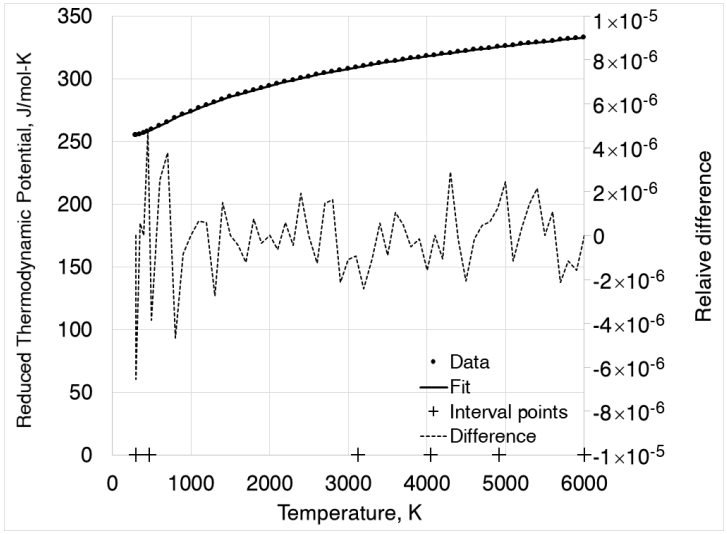
Lead Chloride, Ion ($PbCl^+$) – Cl1Pb1+(g)

MW = 242.7 g/mol – $\Delta_f H^o$(298.15 K) = 745.589 kJ/mol – S^o(298.15 K) = 255.148 J/mol-K – p^o = 0.1 MPa



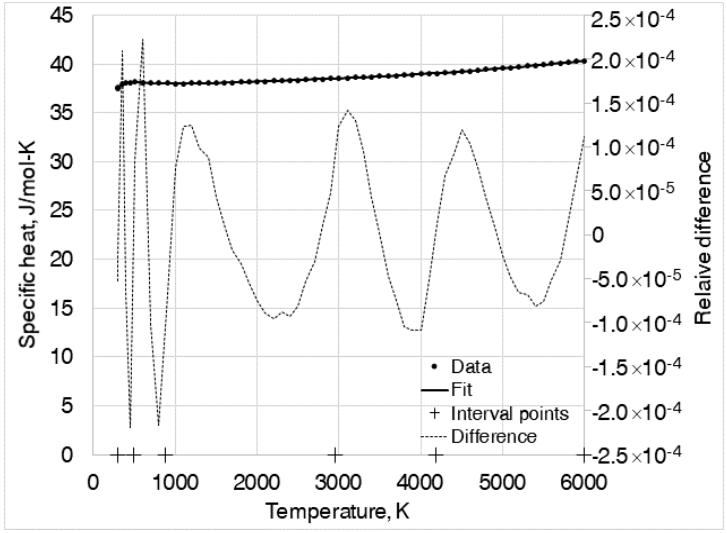


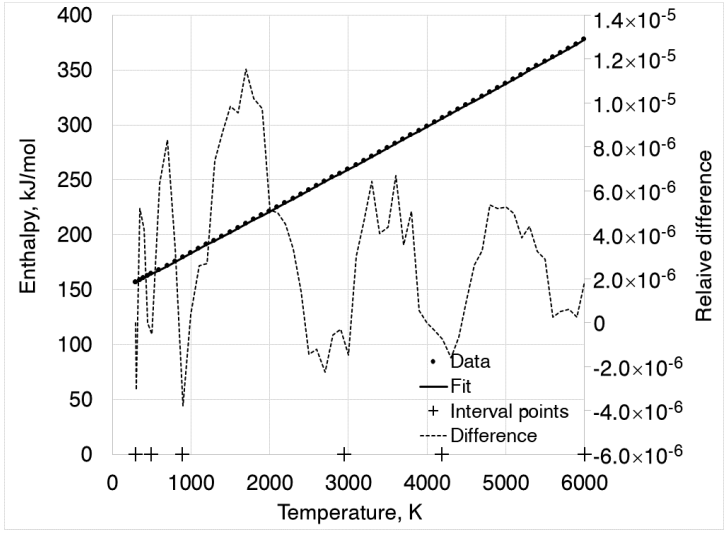


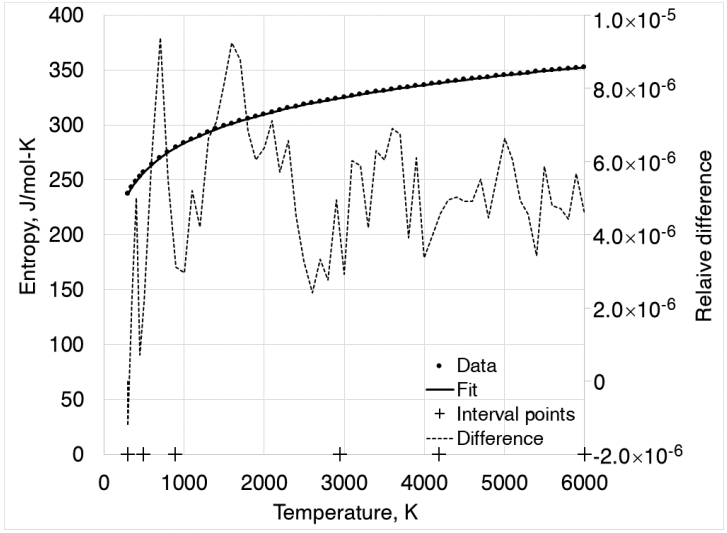


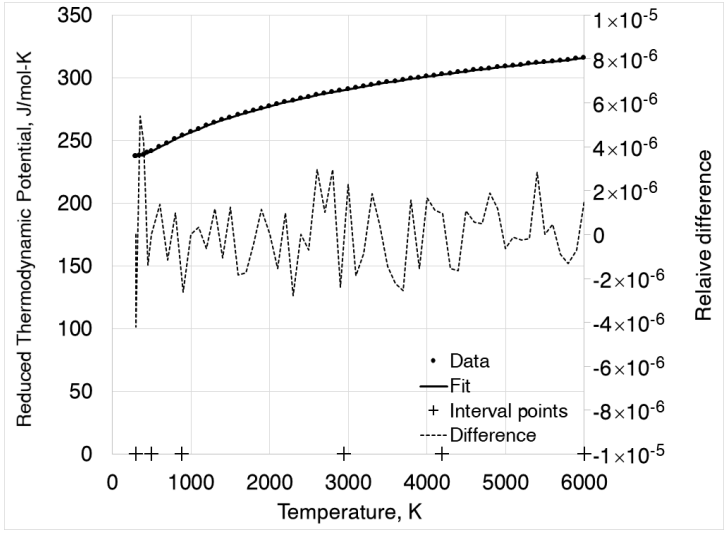
Sulfur Chloride (SCl) – Cl1S1(g)

MW = 67.51 g/mol – $\Delta_f H^o$(298.15 K) = 156.465 kJ/mol – S^o(298.15 K) = 237.324 J/mol-K – p^o = 0.1 MPa



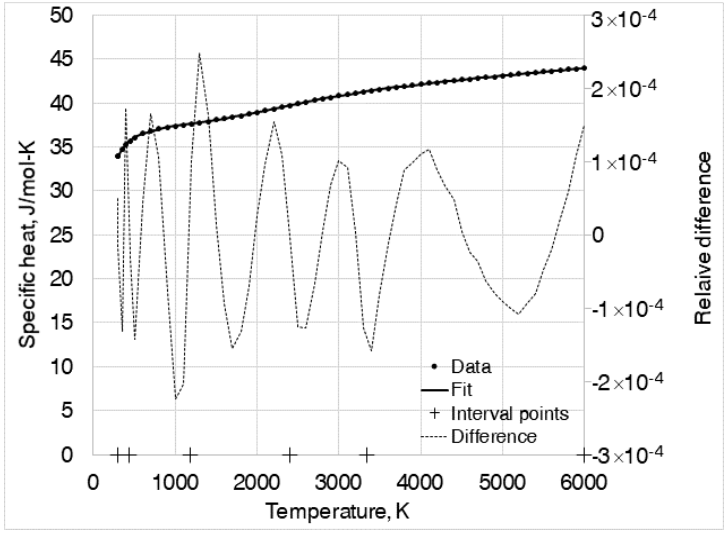


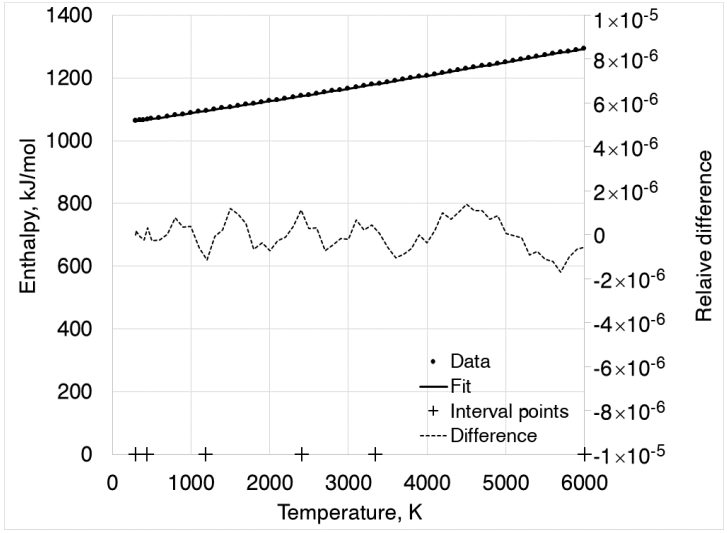


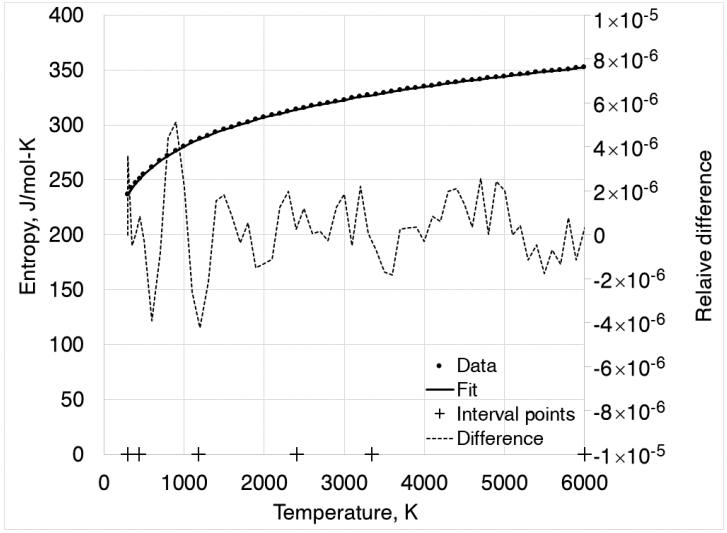


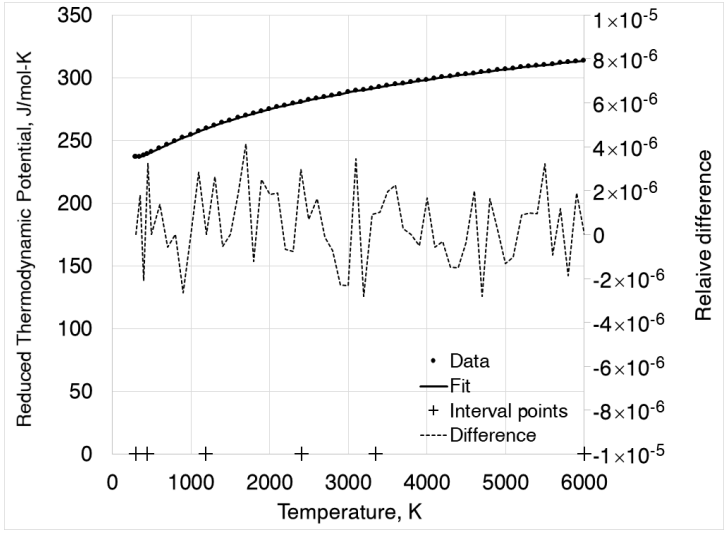
Sulfur Chloride, Ion (SCl^+) – Cl1S1+(g)

MW = 67.51 g/mol – $\Delta_f H^o$(298.15 K) = 1062.037 kJ/mol – S^o(298.15 K) = 236.633 J/mol-K – p^o = 0.1 MPa



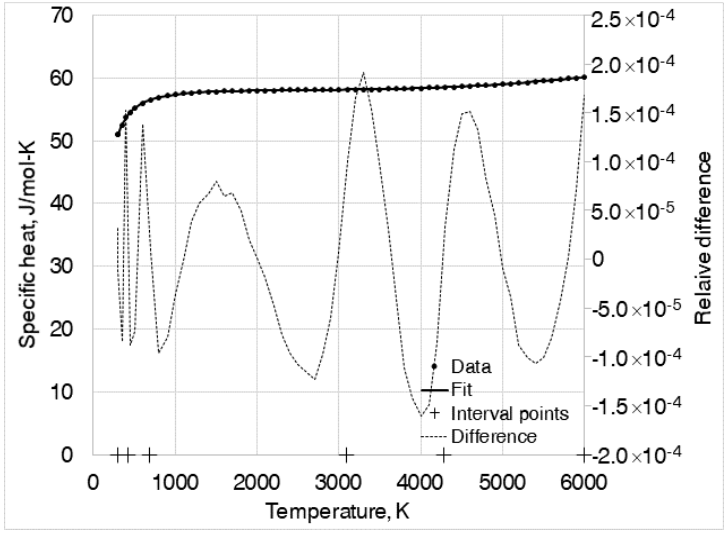


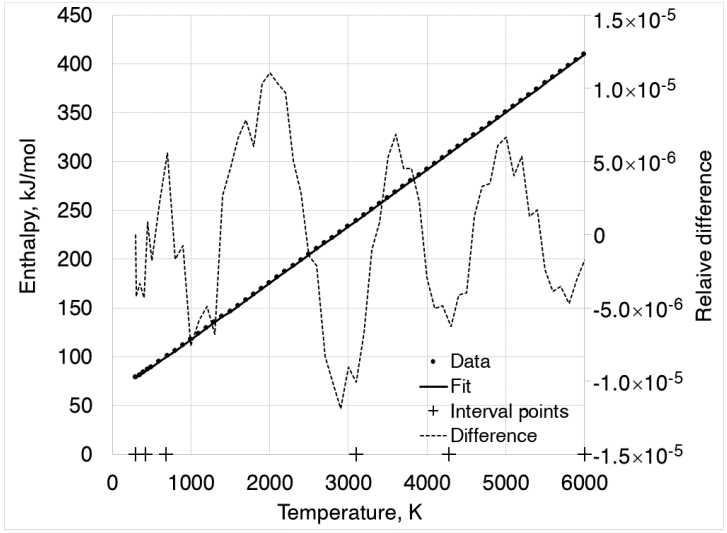


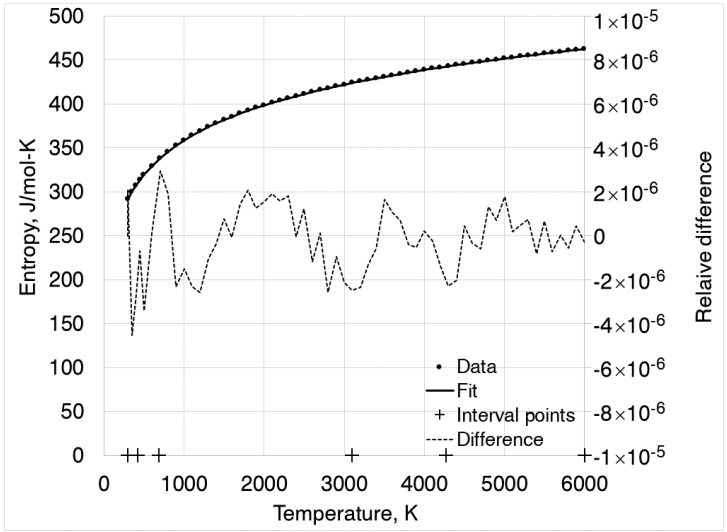


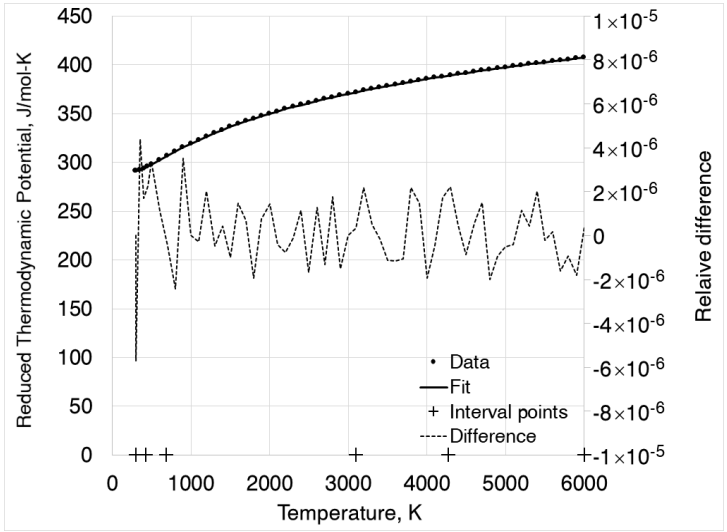
Sulfur Chloride (S_2Cl) – Cl1S2(g)

MW = 99.57 g/mol – $\Delta_f H^o$(298.15 K) = 78.559 kJ/mol – S^o(298.15 K) = 291.686 J/mol-K – p^o = 0.1 MPa



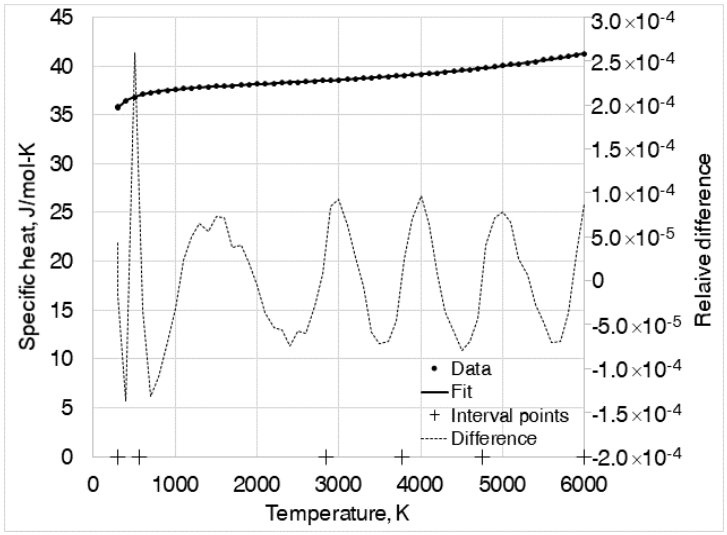


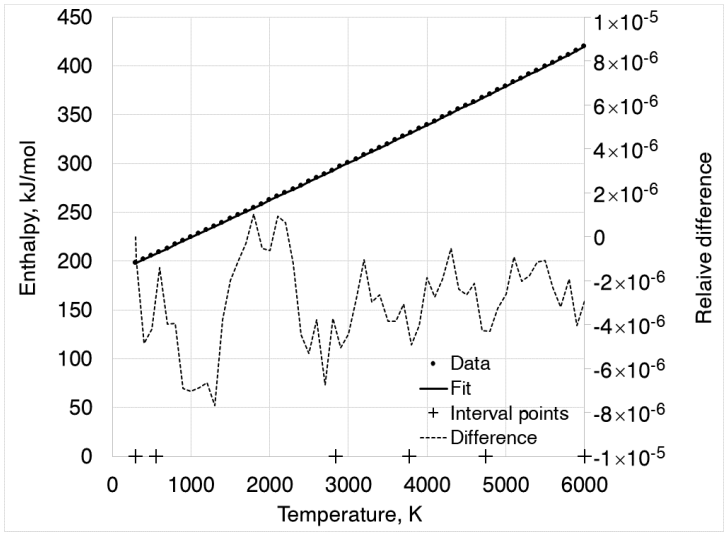


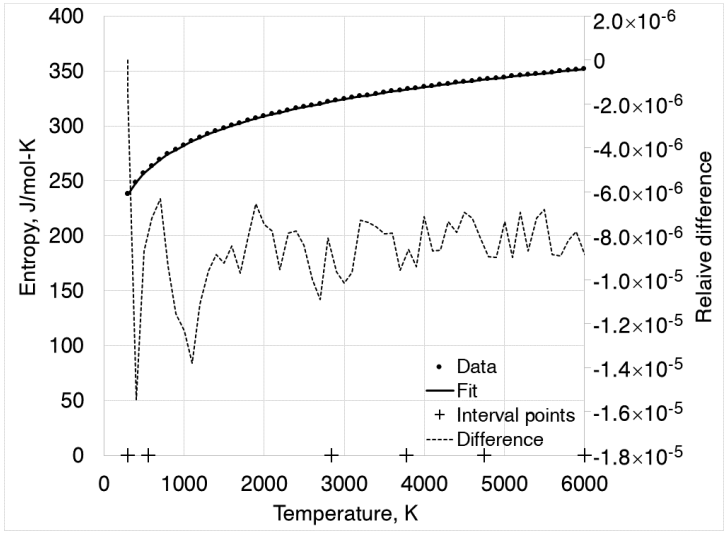


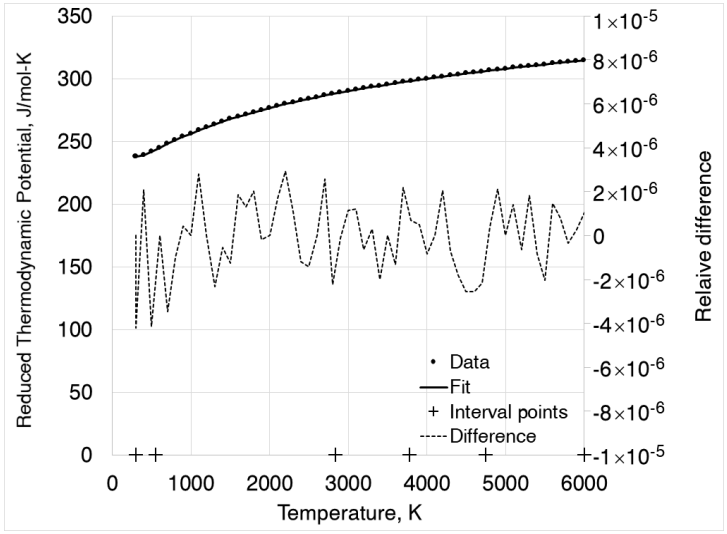
Clorosilylidyne (SiCl) – Cl1Si1(g)

MW = 63.538 g/mol – $\Delta_f H^o$(298.15 K) = 198.322 kJ/mol – S^o(298.15 K) = 237.833 J/mol-K – p^o = 0.1 MPa



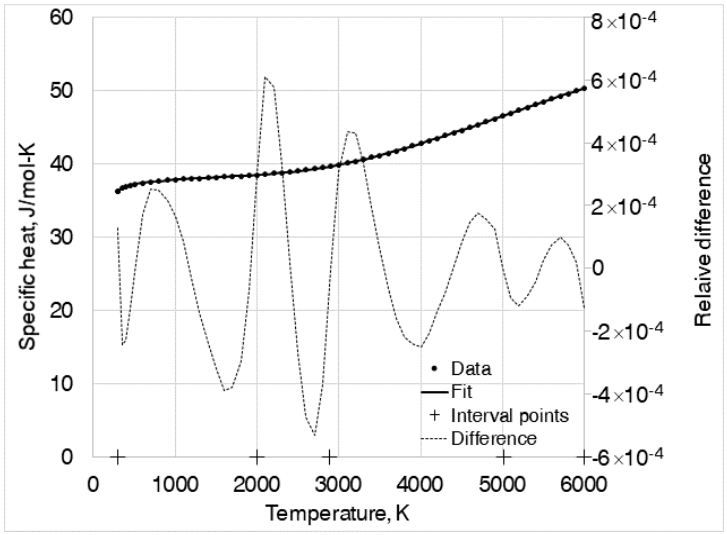


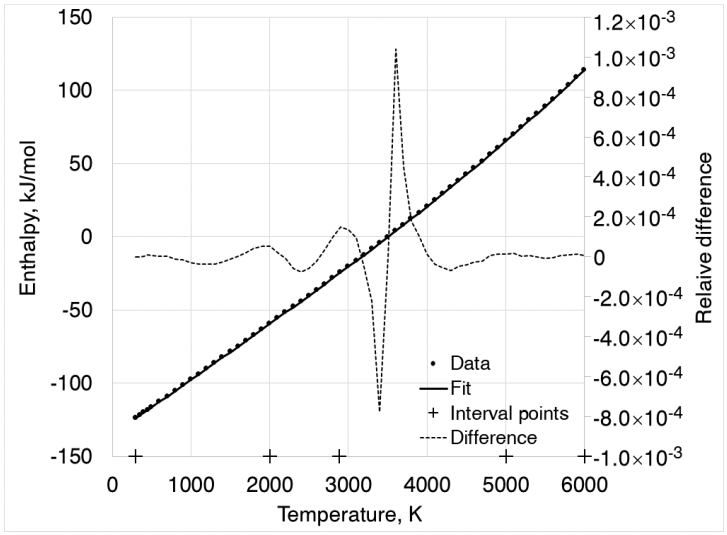


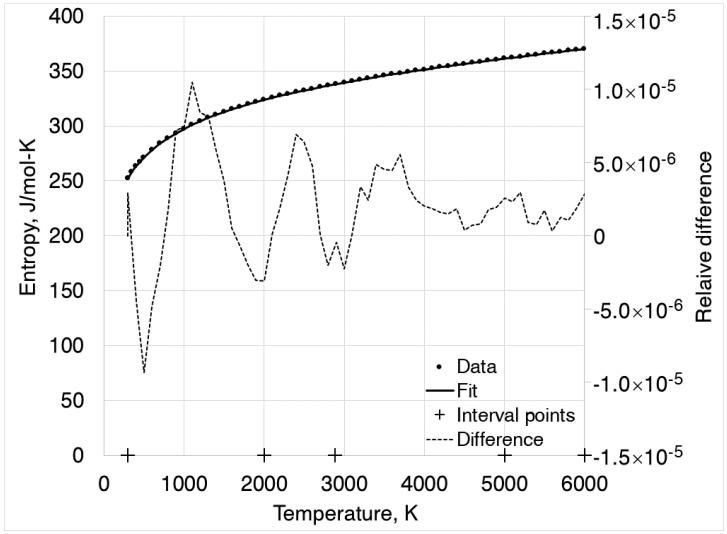


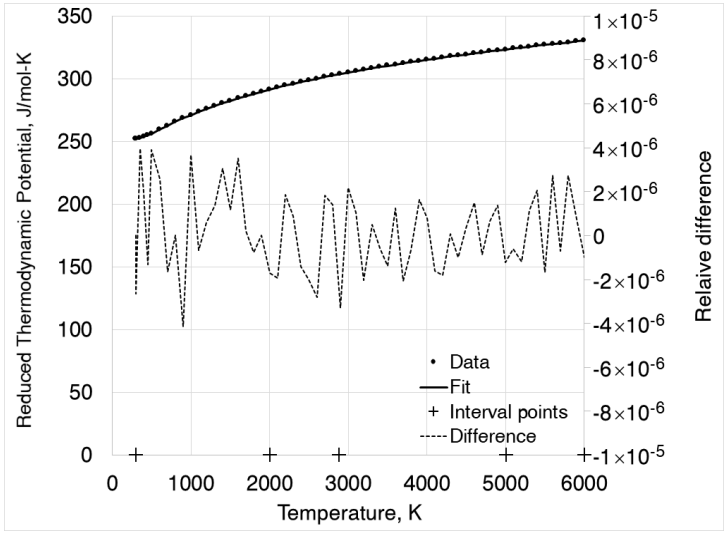
Strontium Chloride (SrCl) – Cl1Sr1(g)

MW = 123.07 g/mol – $\Delta_f H^o$(298.15 K) = -123.846 kJ/mol – S^o(298.15 K) = 252.284 J/mol-K – p^o = 0.1 MPa



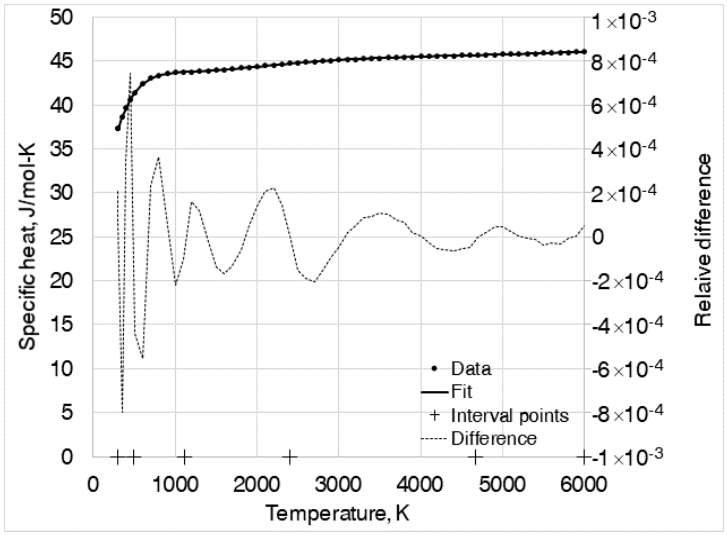


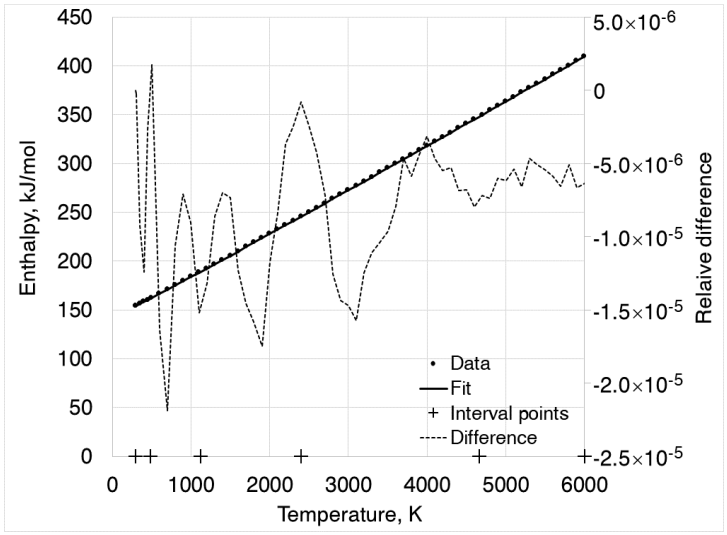


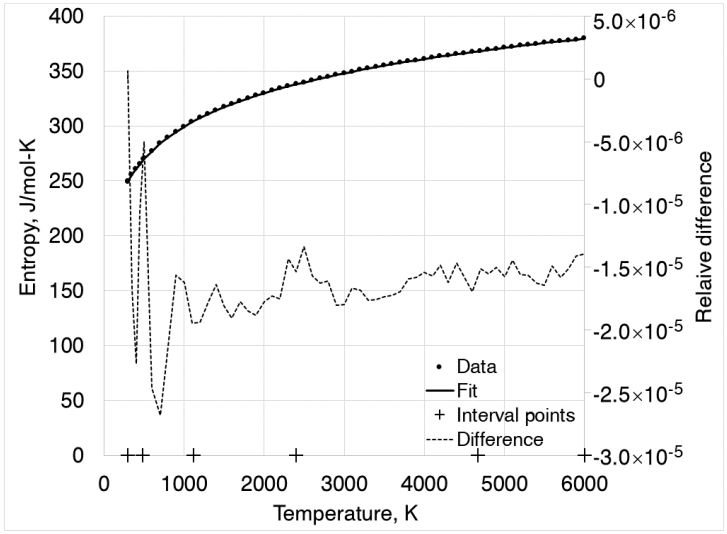


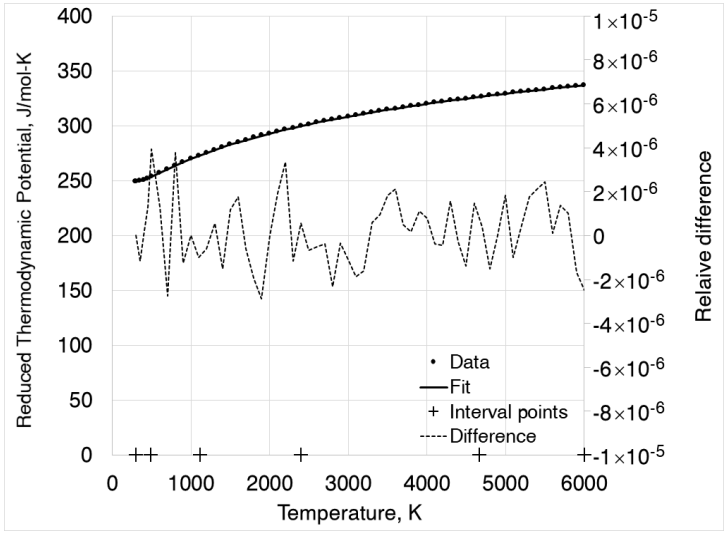
Titanium Chloride (TiCl) – Cl1Ti1(g)

MW = 83.33 g/mol – $\Delta_f H^o$(298.15 K) = 154.39 kJ/mol – S^o(298.15 K) = 249.214 J/mol-K – p^o = 0.1 MPa



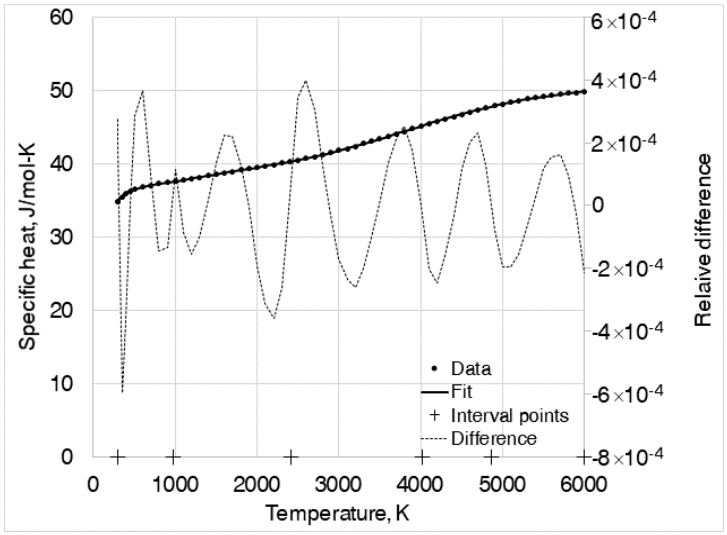


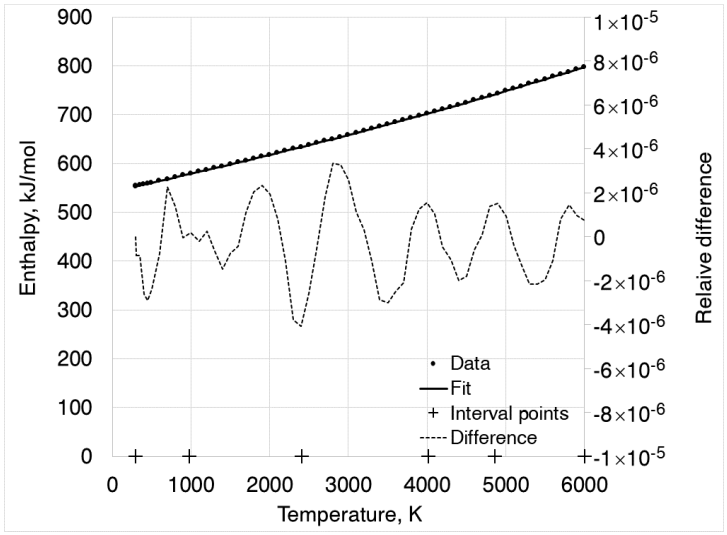


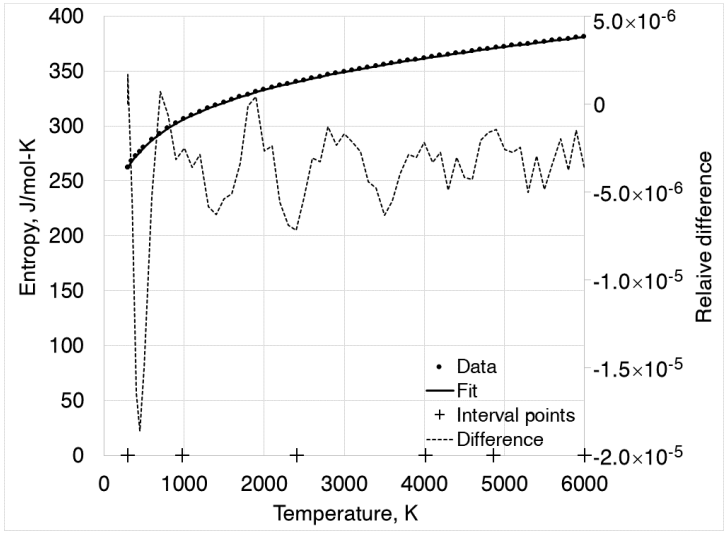


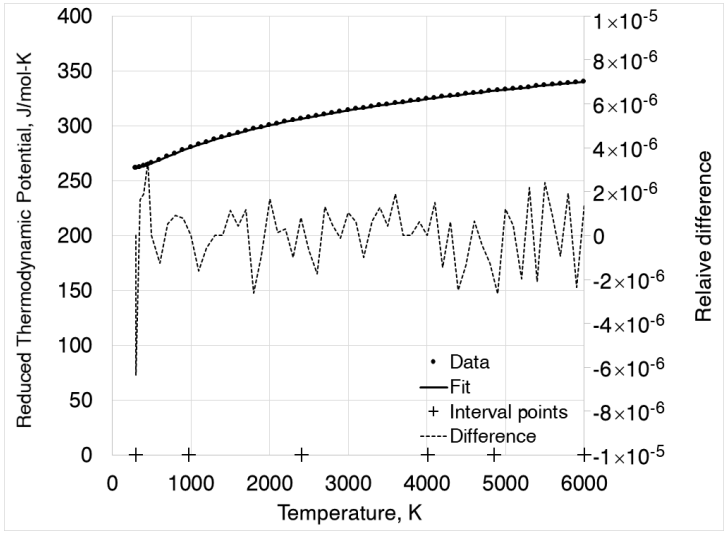
Tungsten Chloride (WCl) – Cl1W1(g)

MW = 219.3 g/mol – $\Delta_f H^o$(298.15 K) = 553.543 kJ/mol – S^o(298.15 K) = 261.895 J/mol-K – p^o = 0.1 MPa



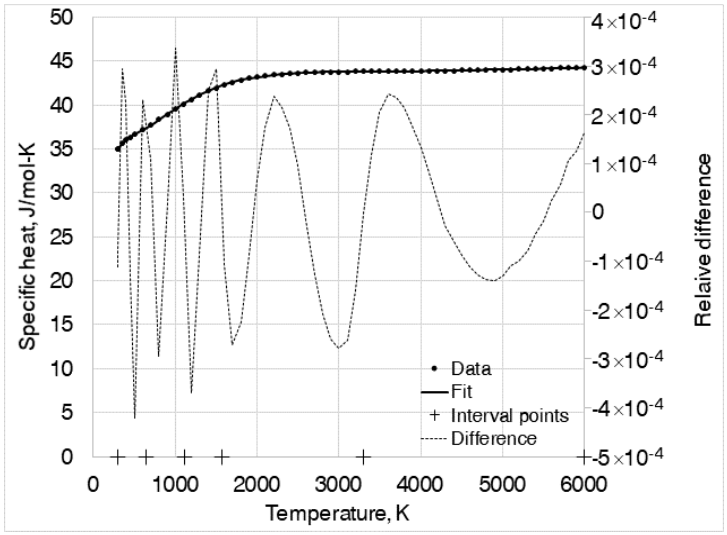


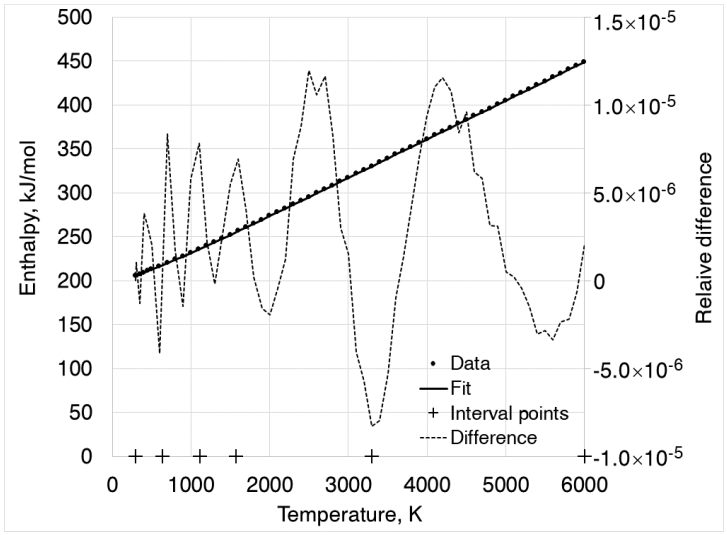


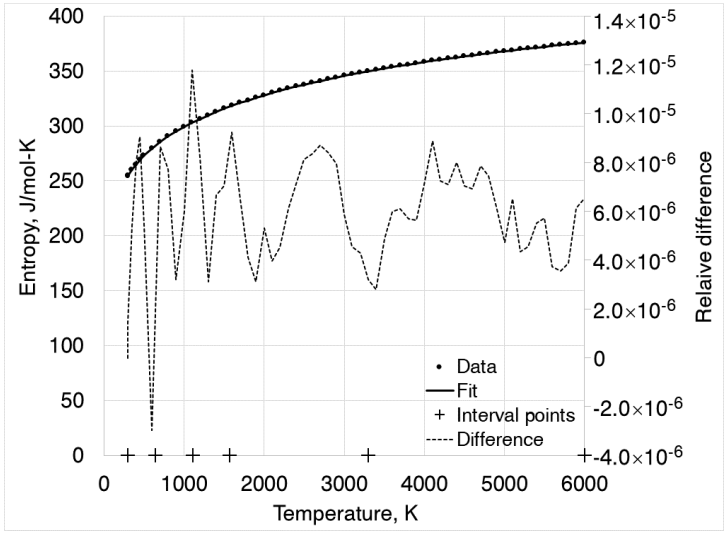


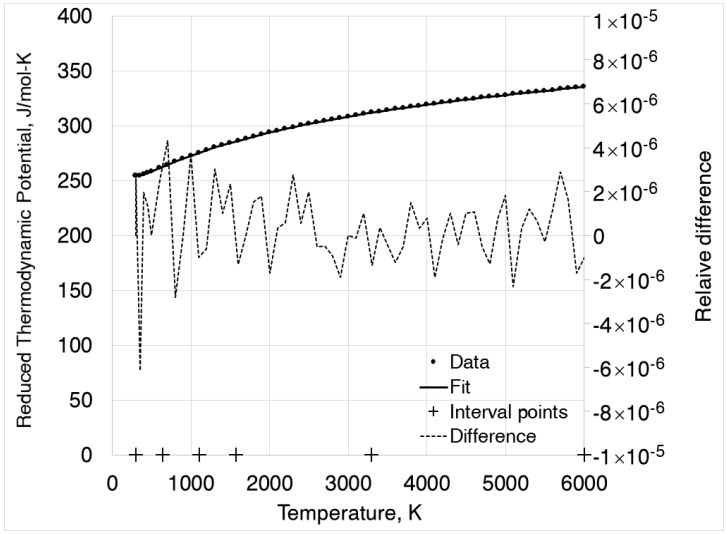
Zirconium Chloride (ZrCl) – Cl1Zr1(g)

MW = 126.67 g/mol – $\Delta_f H^o$(298.15 K) = 205.434 kJ/mol – S^o(298.15 K) = 254.309 J/mol-K – p^o = 0.1 MPa



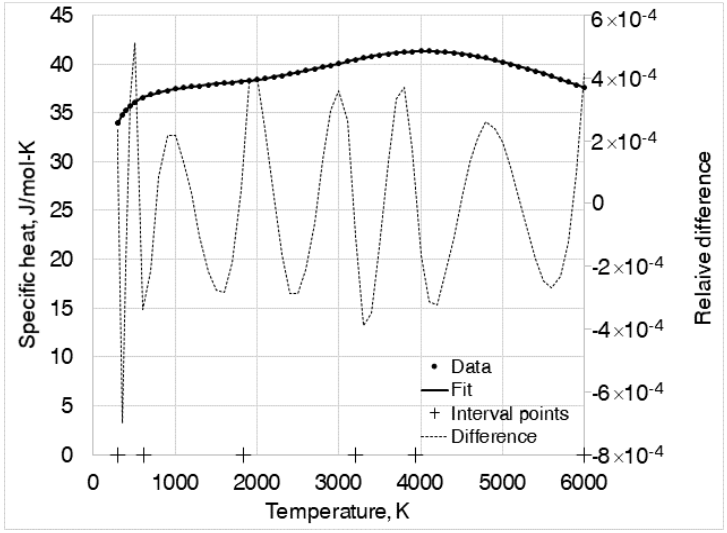


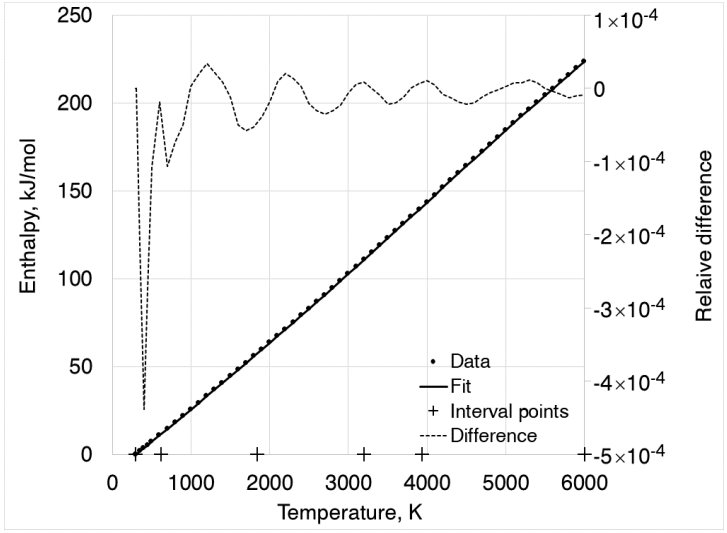


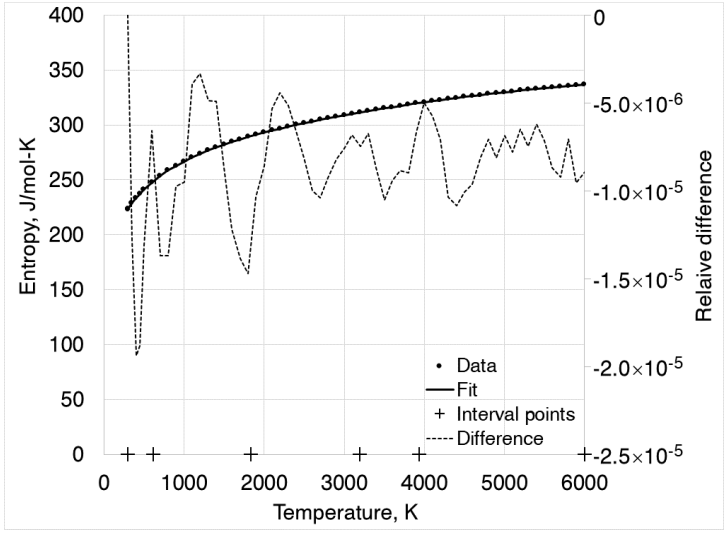


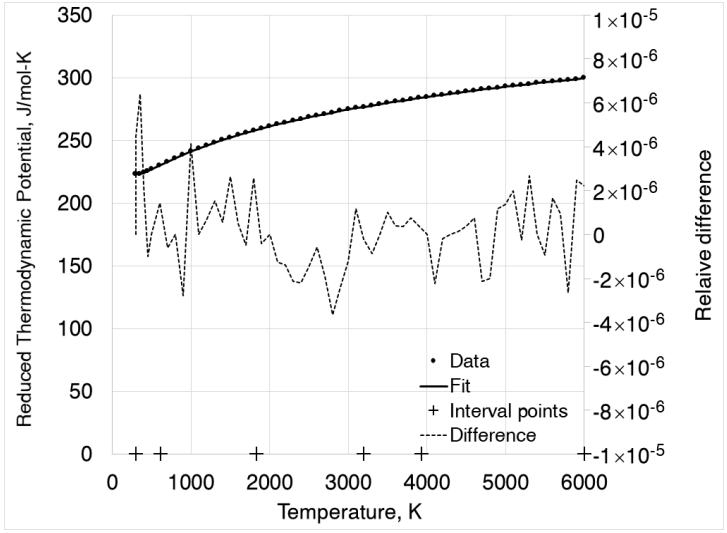
Chlorine (Cl_2) – Cl2(ref)

MW = 70.906 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 223.079 J/mol-K – p^o = 0.1 MPa



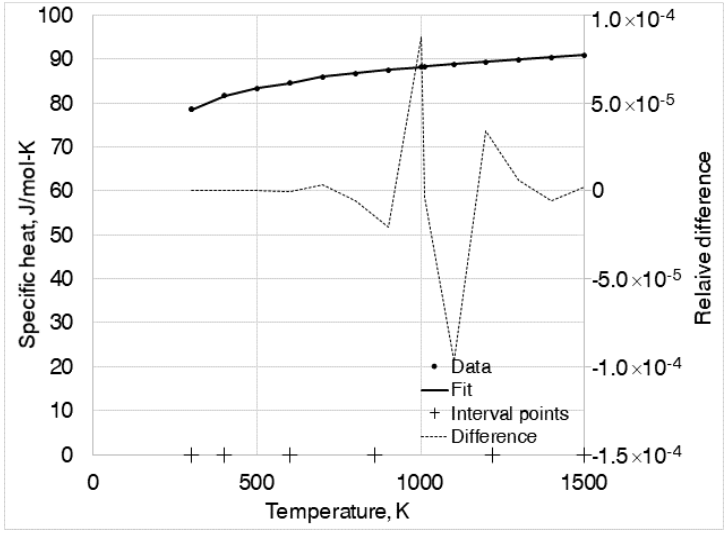


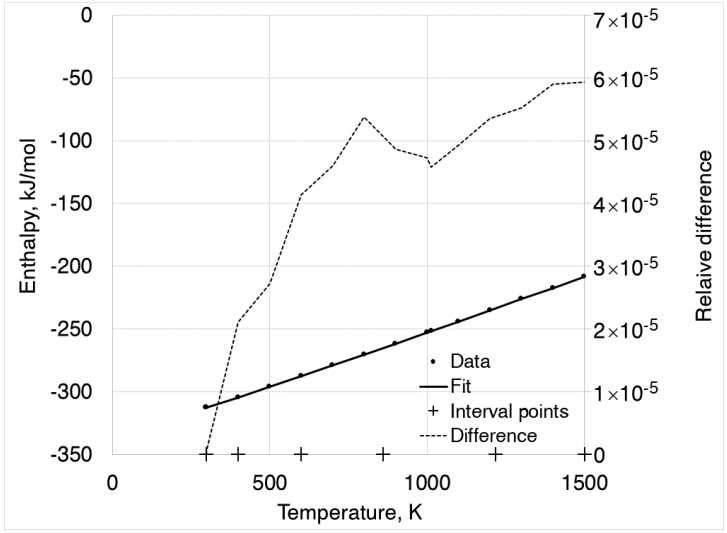


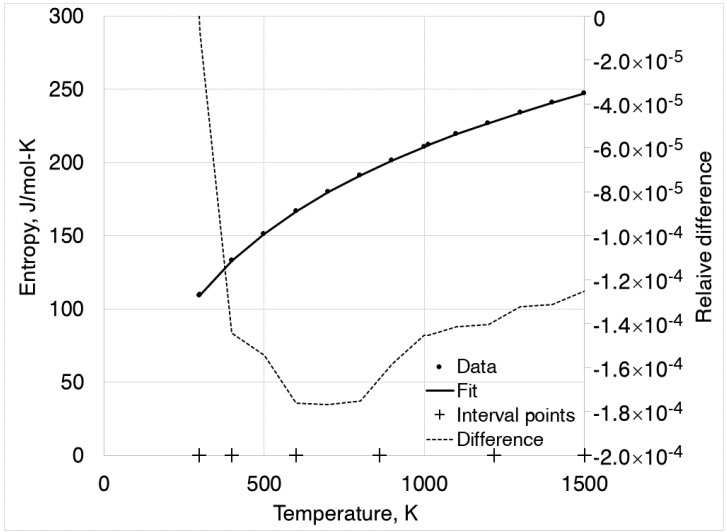


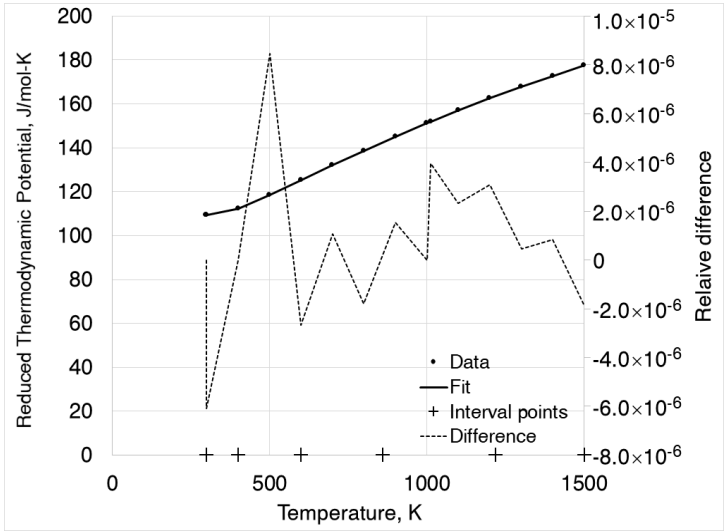
Cobalt Chloride ($CoCl_2$) – Cl2Co1(cr)

MW = 129.837 g/mol – $\Delta_f H^o$(298.15 K) = -312.545 kJ/mol – S^o(298.15 K) = 109.266 J/mol-K – p^o = 0.1 MPa



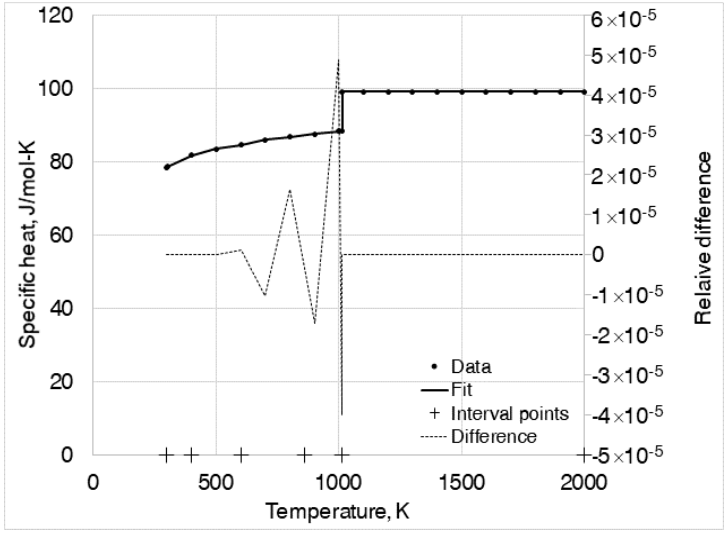


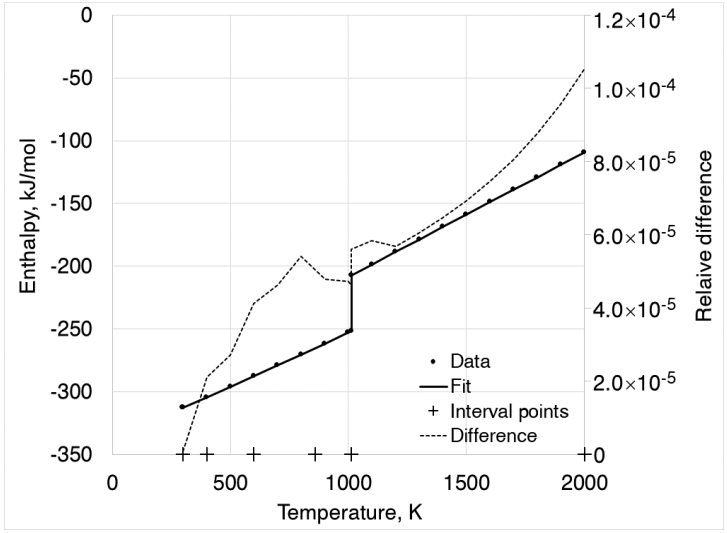


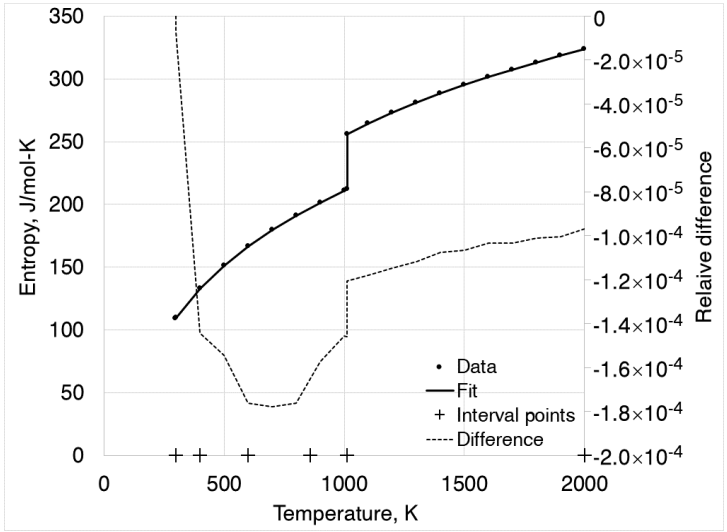


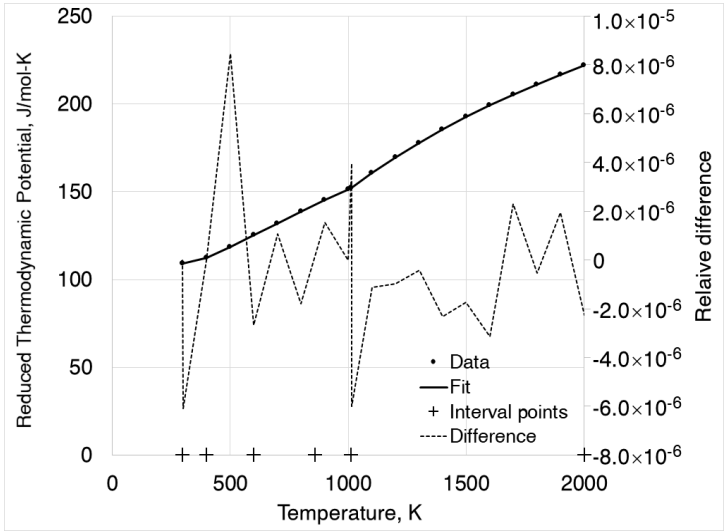
Cobalt Chloride ($CoCl_2$) – Cl2Co1(cr,l)

MW = 129.837 g/mol – $\Delta_f H^o$(298.15 K) = -312.545 kJ/mol – S^o(298.15 K) = 109.266 J/mol-K – p^o = 0.1 MPa



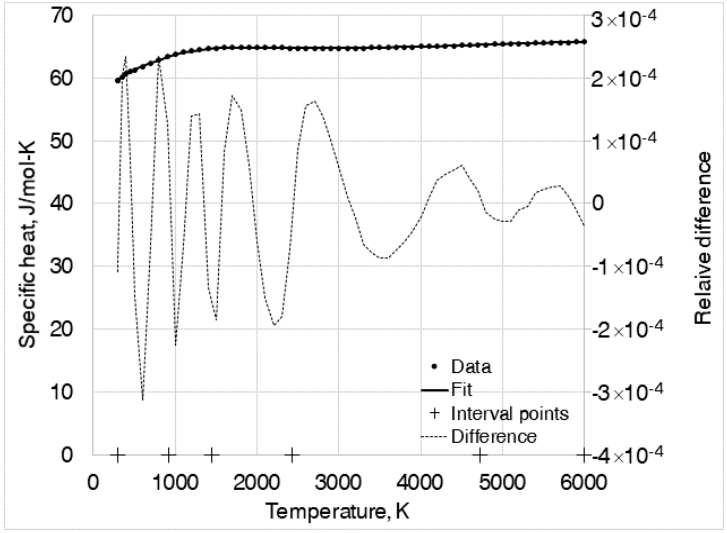


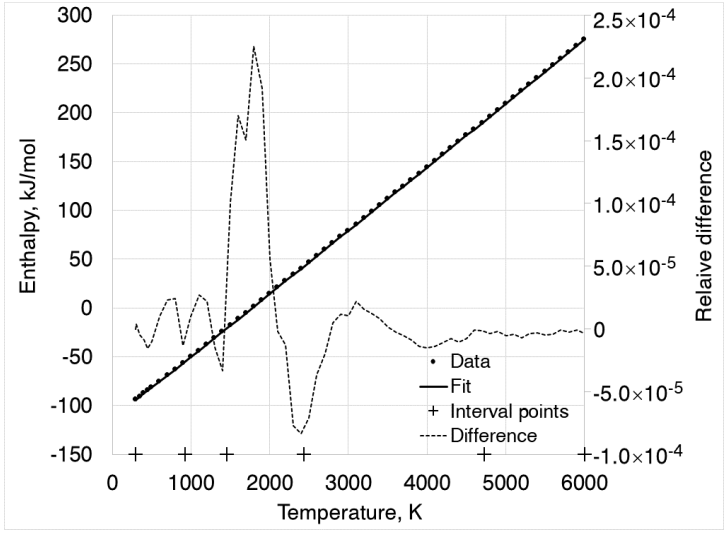


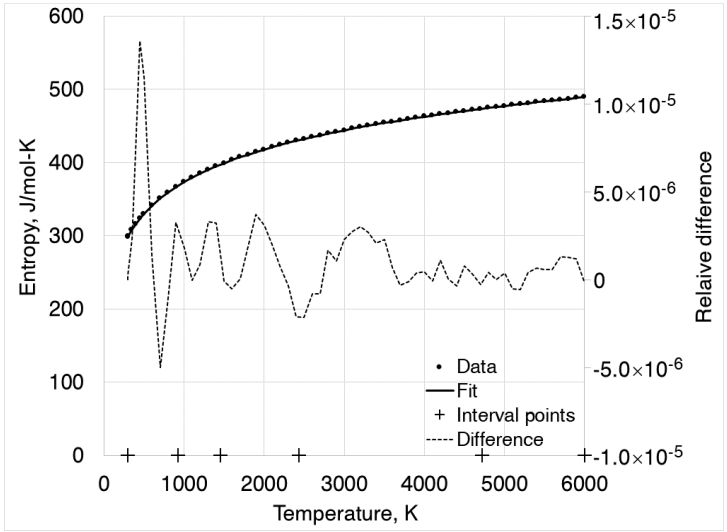


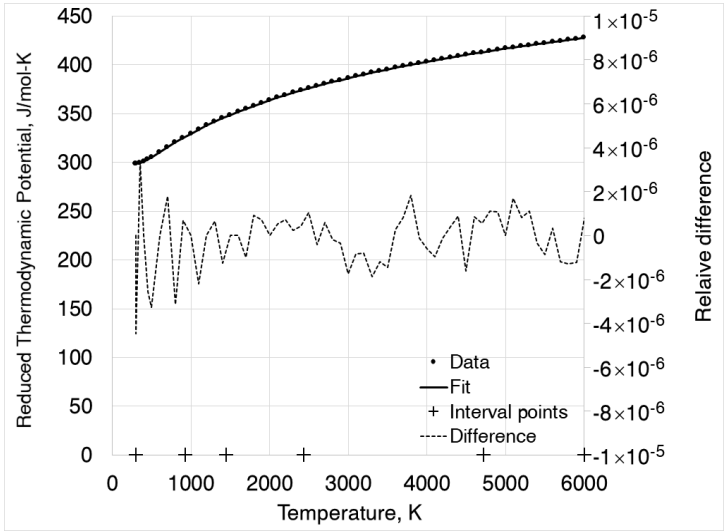
Cobalt Chloride ($CoCl_2$) – Cl2Co1(g)

MW = 129.837 g/mol – $\Delta_f H^o$(298.15 K) = -93.722 kJ/mol – S^o(298.15 K) = 298.5 J/mol-K – p^o = 0.1 MPa



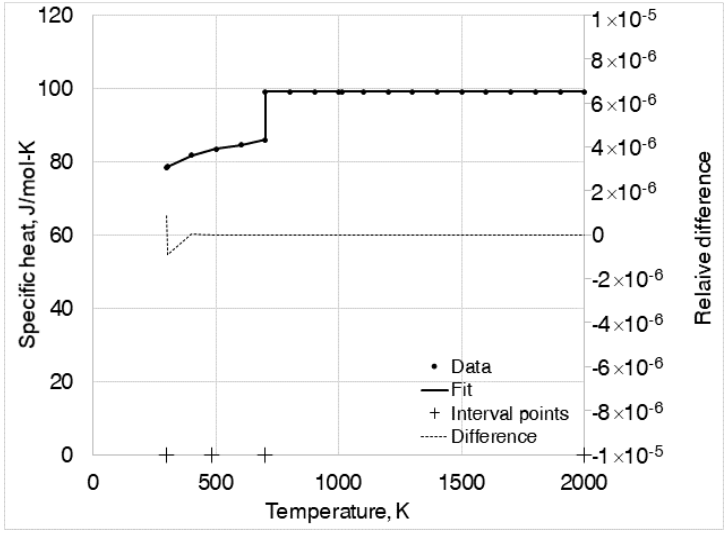


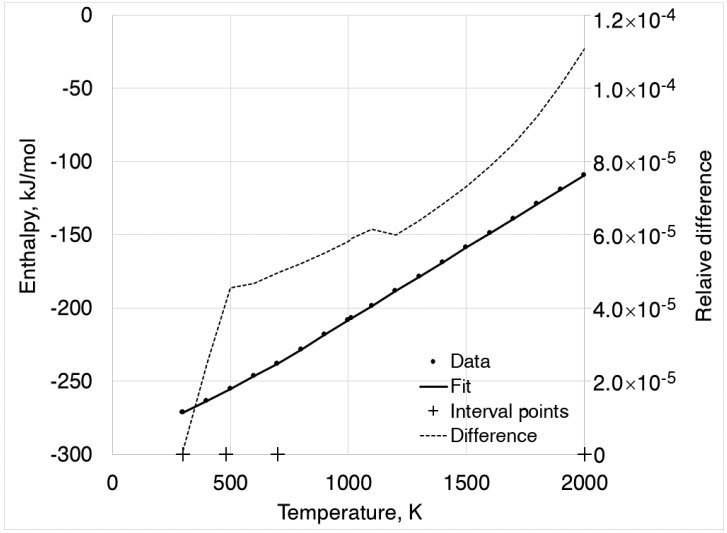


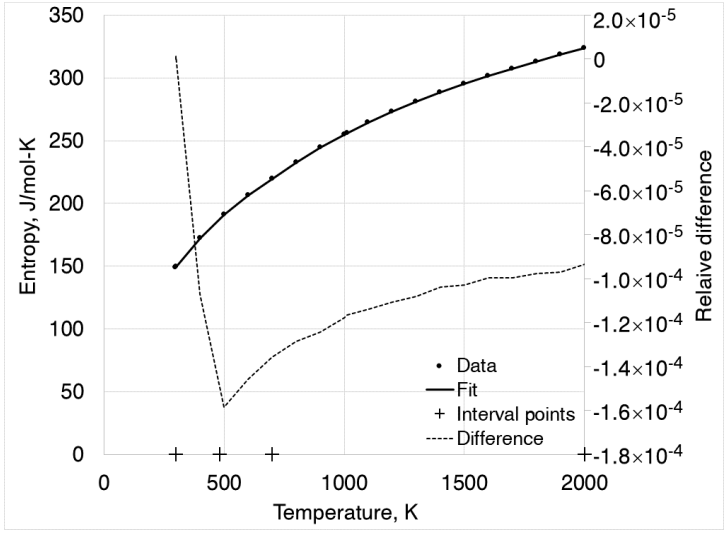


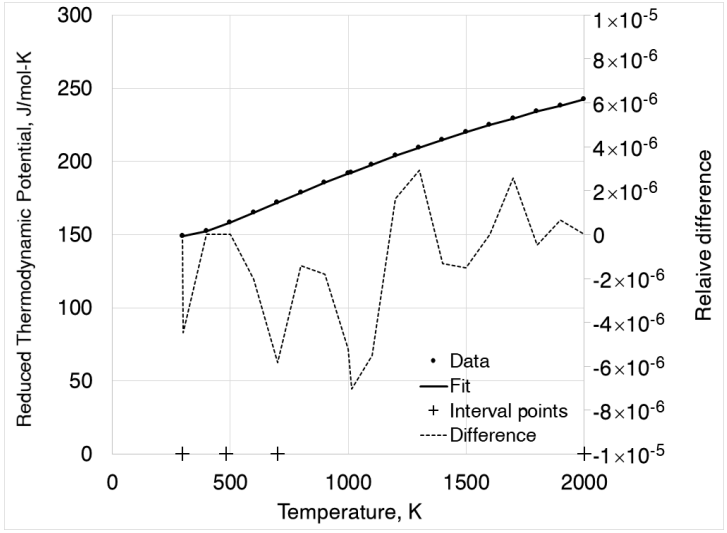
Cobalt Chloride ($CoCl_2$) – Cl2Co1(l)

MW = 129.837 g/mol – $\Delta_f H^o$(298.15 K) = -271.514 kJ/mol – S^o(298.15 K) = 149.021 J/mol-K – p^o = 0.1 MPa



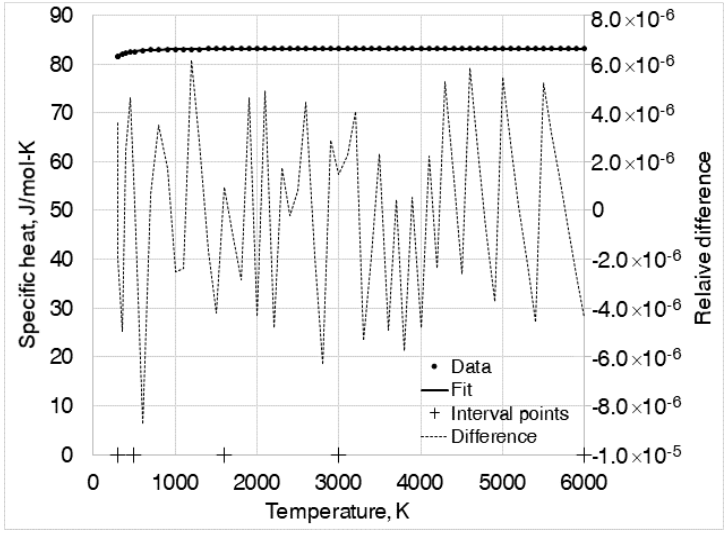


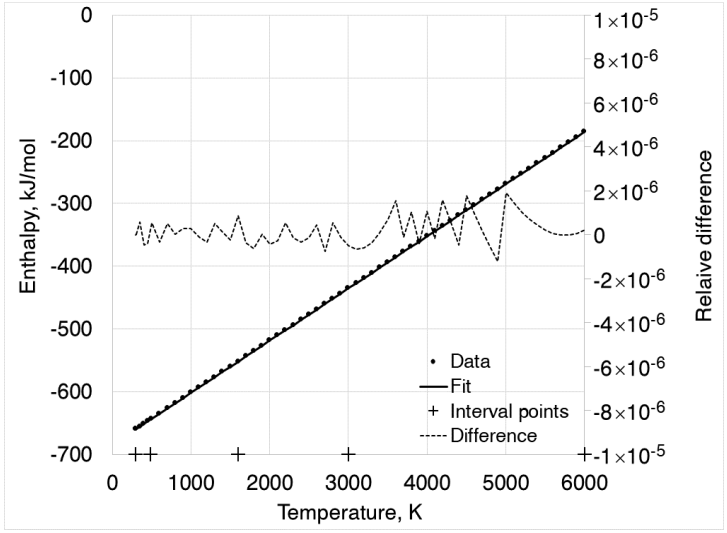


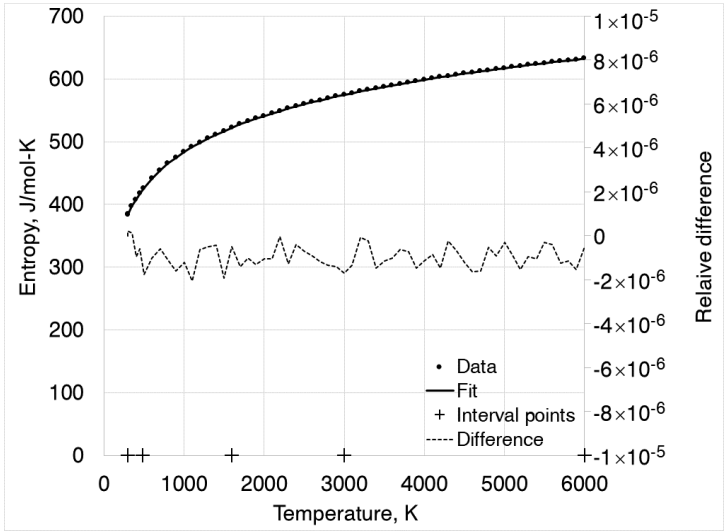


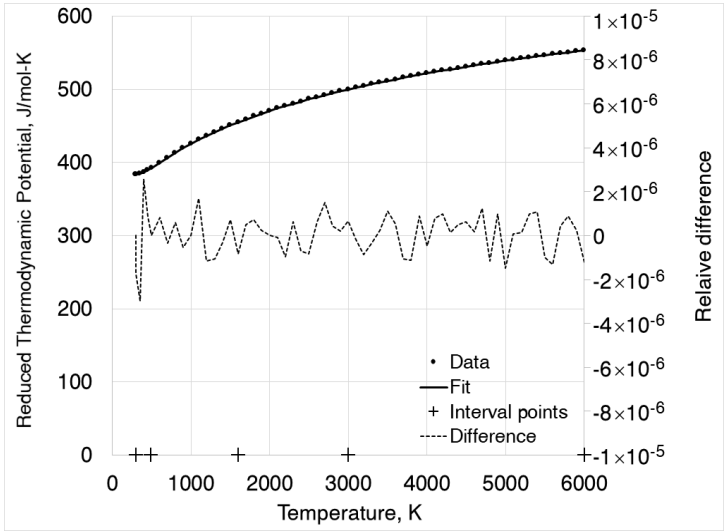
Cesium Chloride ($(CsCl)_2$) – Cl2Cs2(g)

MW = 336.717 g/mol – $\Delta_f H^o$(298.15 K) = -659.817 kJ/mol – S^o(298.15 K) = 383.445 J/mol-K – p^o = 0.1 MPa



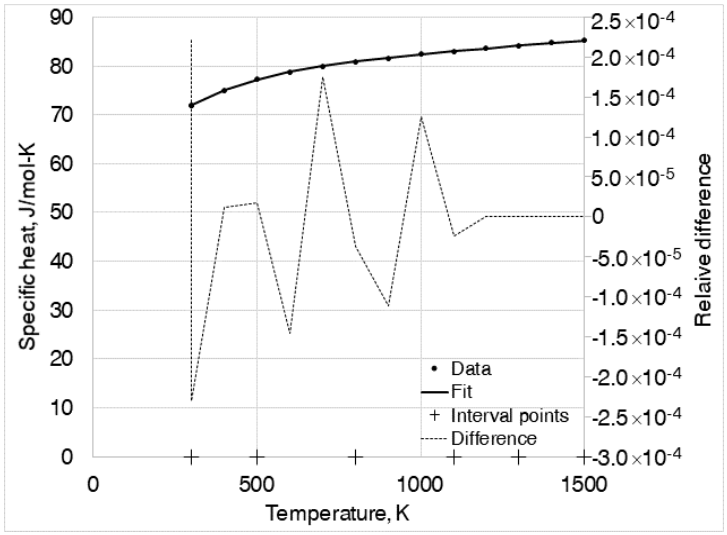


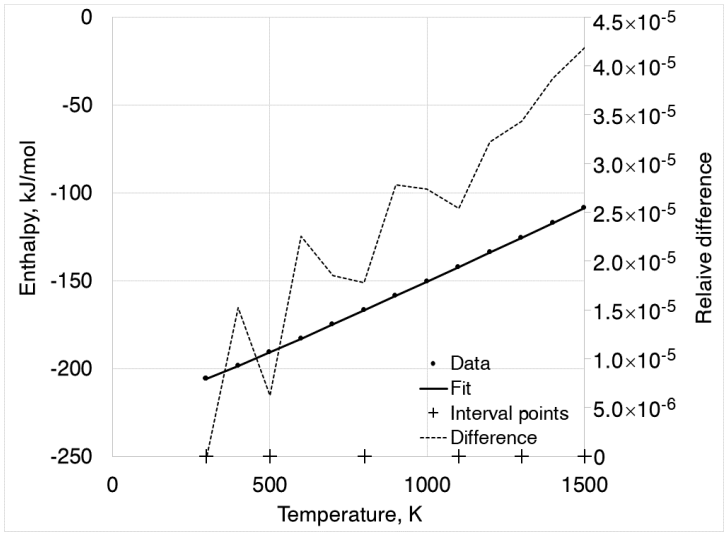


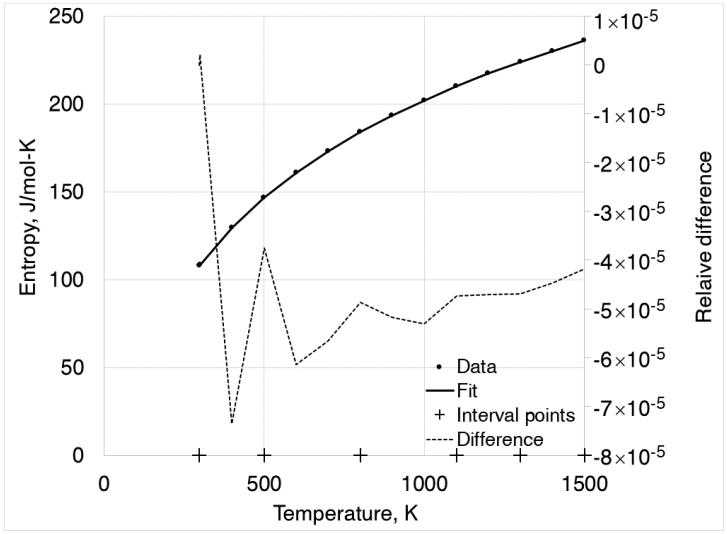


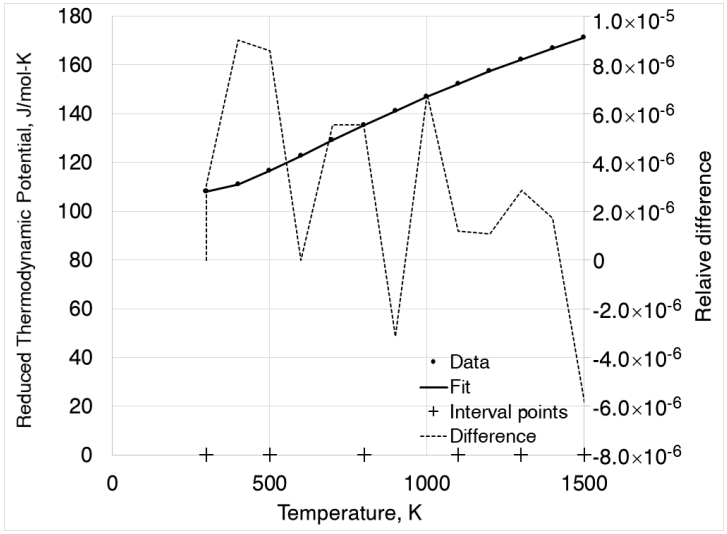
Copper Chloride ($CuCl_2$) – Cl2Cu1(cr)

MW = 134.452 g/mol – $\Delta_f H^o$(298.15 K) = -205.853 kJ/mol – S^o(298.15 K) = 108.085 J/mol-K – p^o = 0.1 MPa



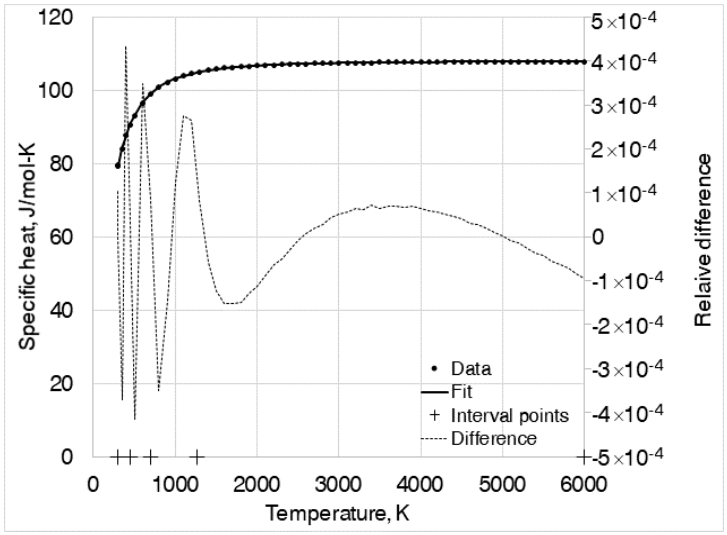


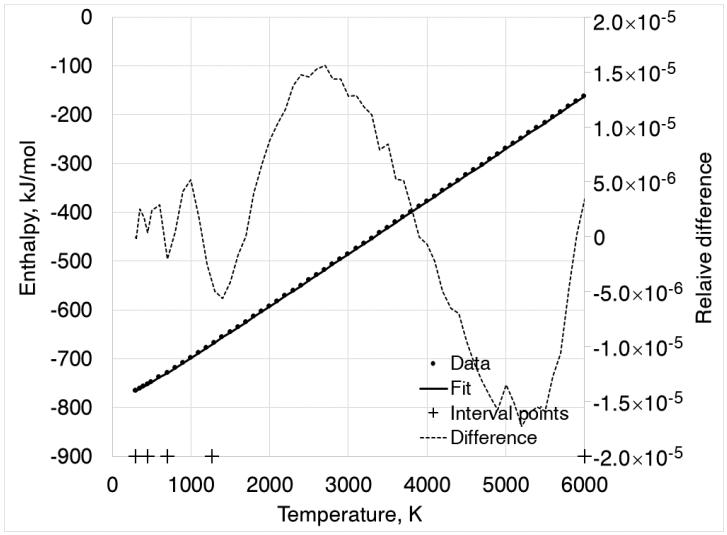


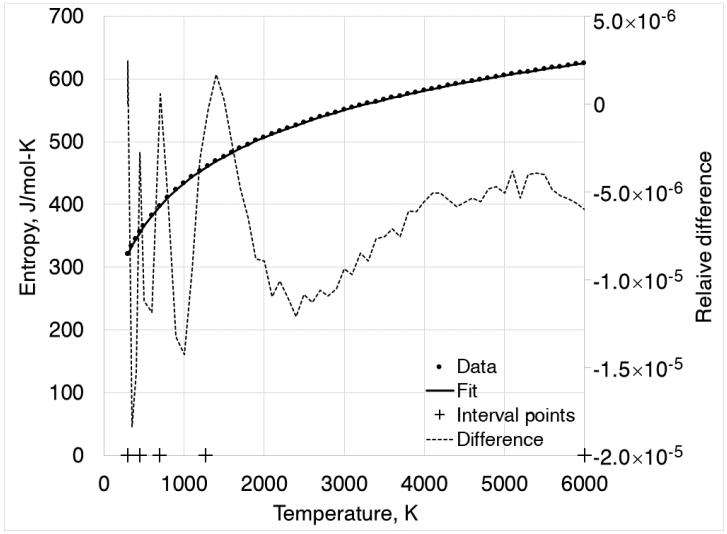


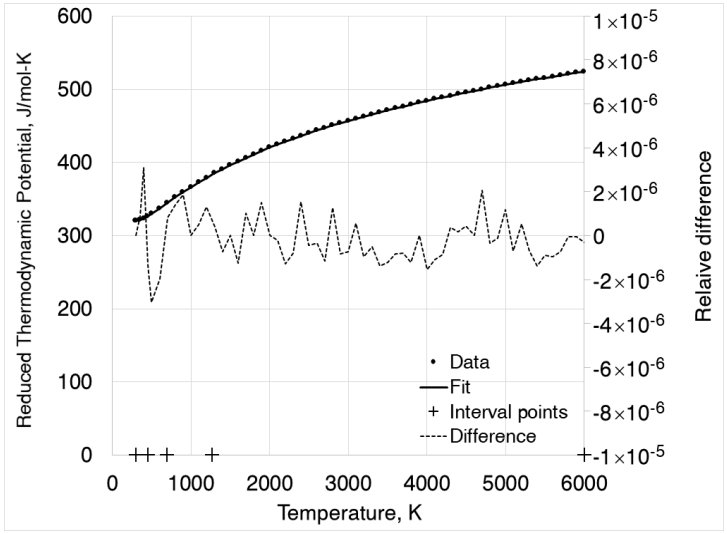
Phosphoryl Chloride Fluoride ($POCl_2F$) – Cl2F1O1P1(g)

MW = 136.878 g/mol – $\Delta_f H^o$(298.15 K) = -765.672 kJ/mol – S^o(298.15 K) = 320.381 J/mol-K – p^o = 0.1 MPa



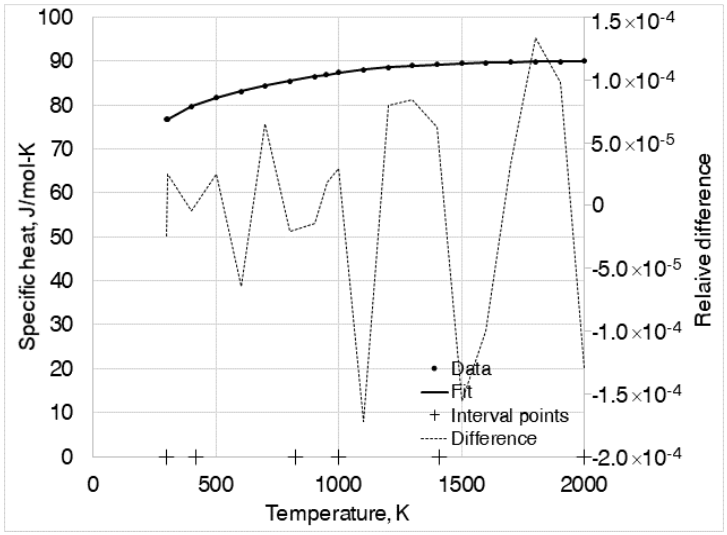


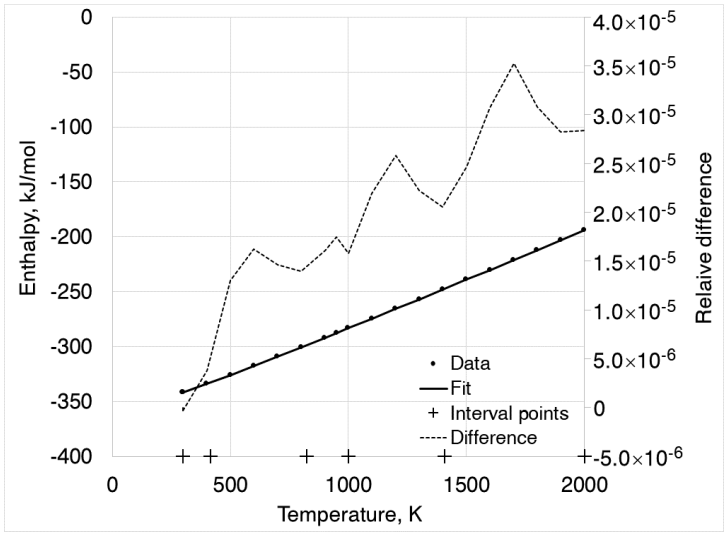


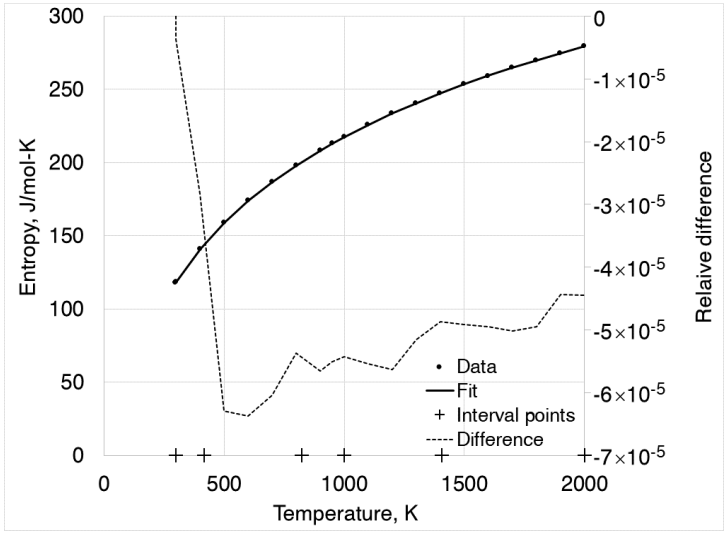


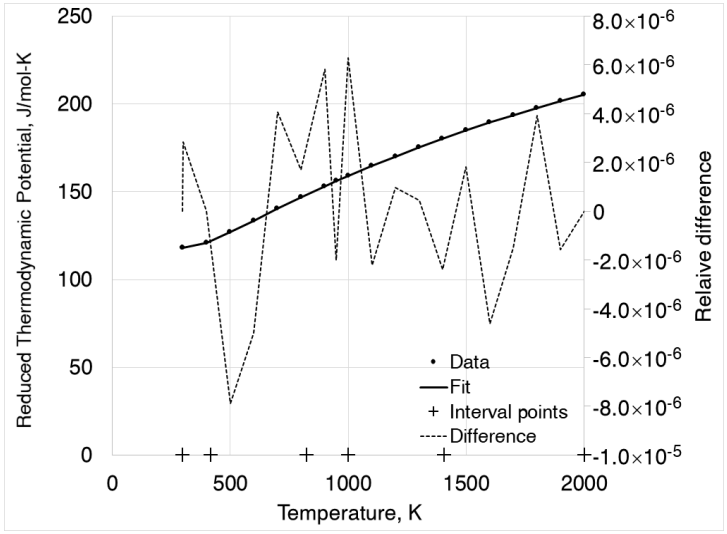
Iron Chloride ($FeCl_2$) – Cl2Fe1(cr)

MW = 126.753 g/mol – $\Delta_f H^o$(298.15 K) = -341.833 kJ/mol – S^o(298.15 K) = 117.947 J/mol-K – p^o = 0.1 MPa



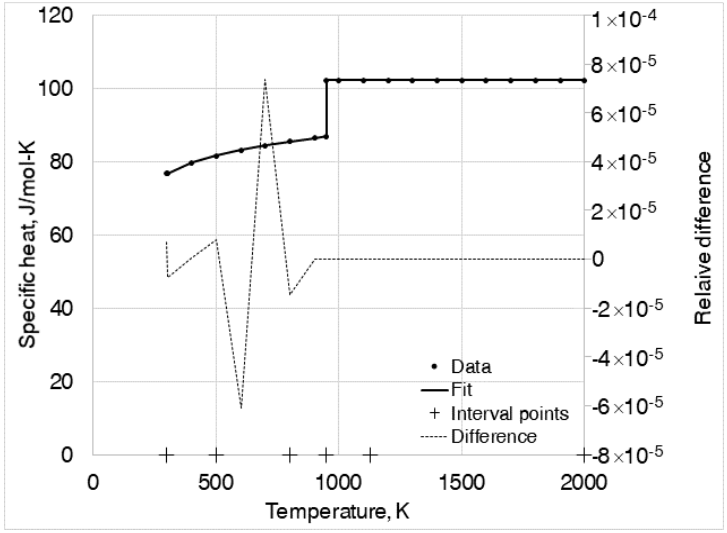


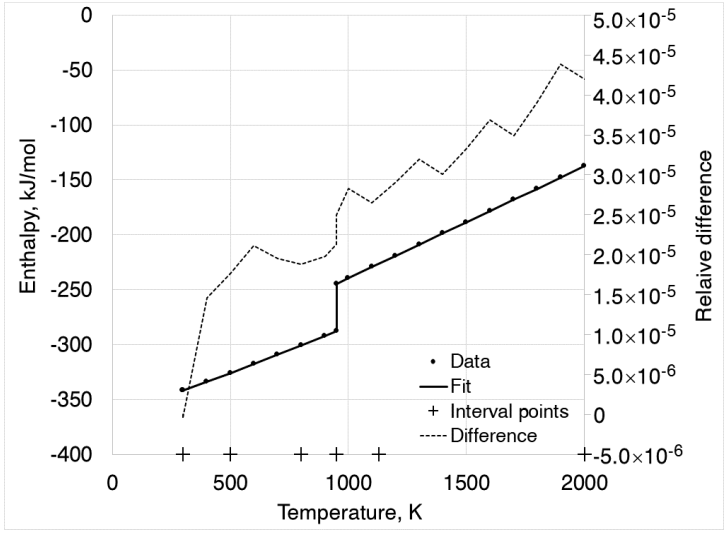


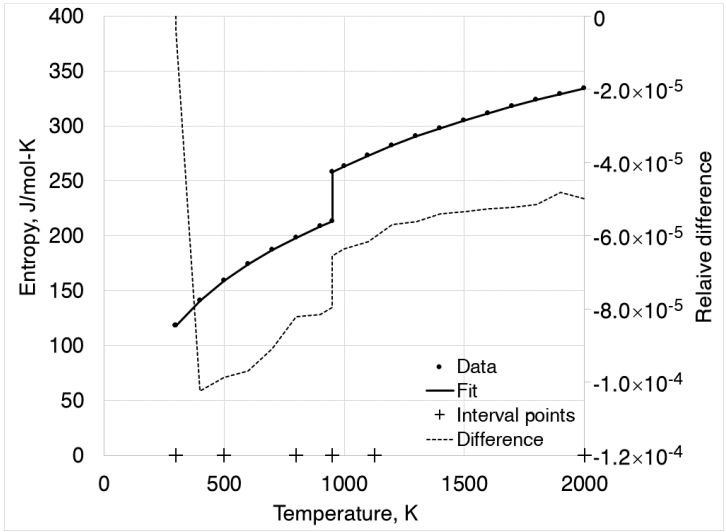


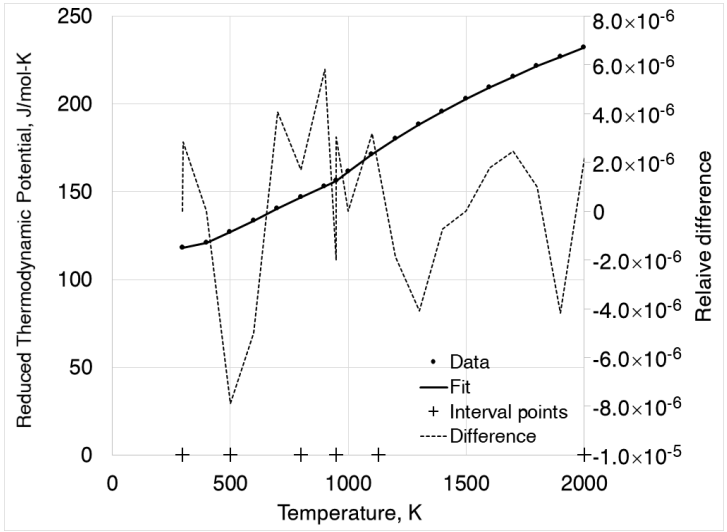
Iron Chloride ($FeCl_2$) – Cl2Fe1(cr,l)

MW = 126.753 g/mol – $\Delta_f H^o$(298.15 K) = -341.833 kJ/mol – S^o(298.15 K) = 117.947 J/mol-K – p^o = 0.1 MPa



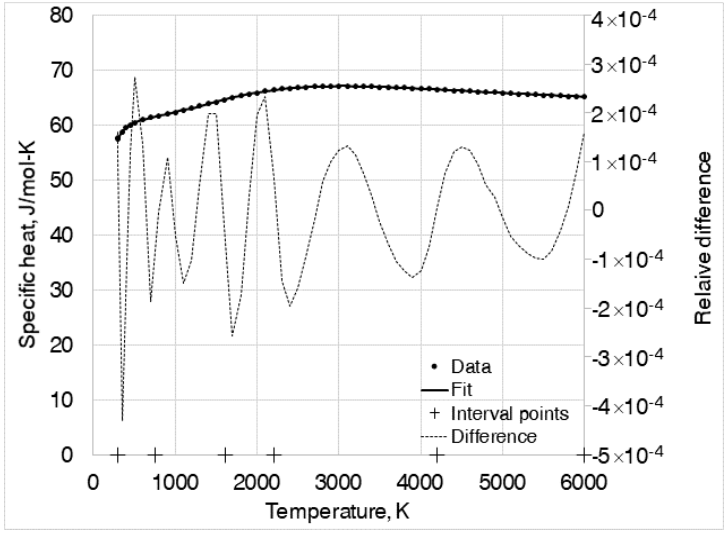


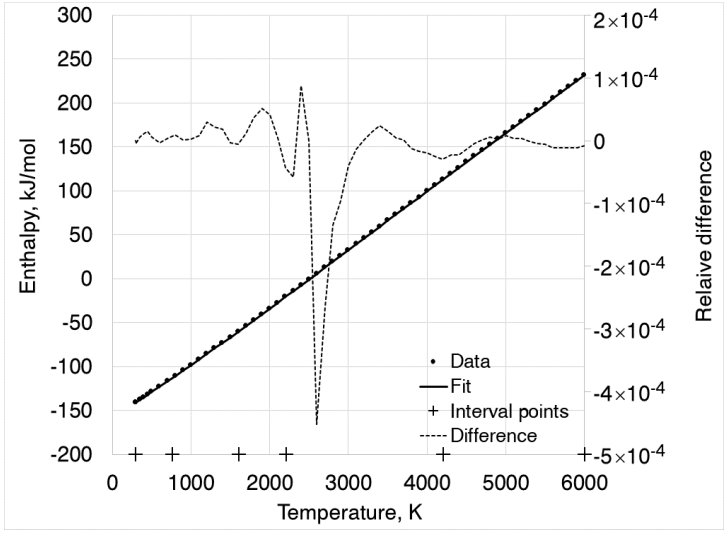


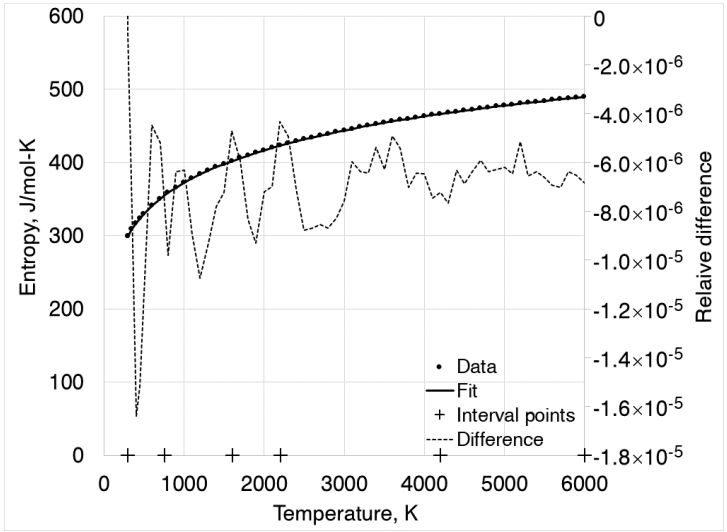


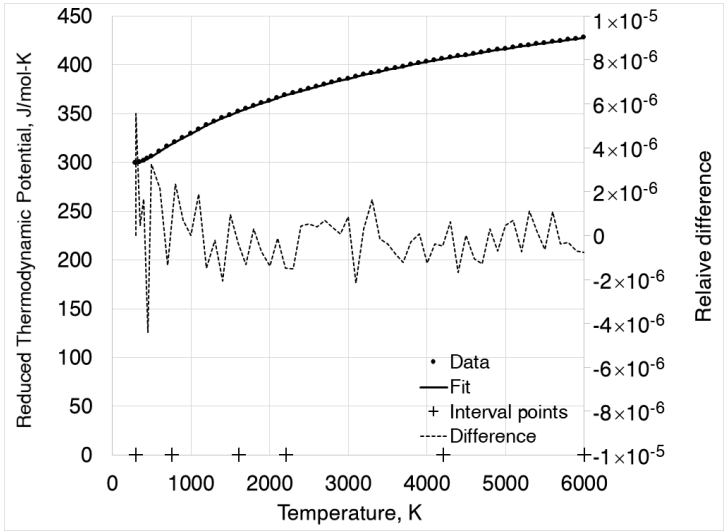
Iron Chloride ($FeCl_2$) – Cl2Fe1(g)

MW = 126.753 g/mol – $\Delta_f H^o$(298.15 K) = -141.001 kJ/mol – S^o(298.15 K) = 299.283 J/mol-K – p^o = 0.1 MPa



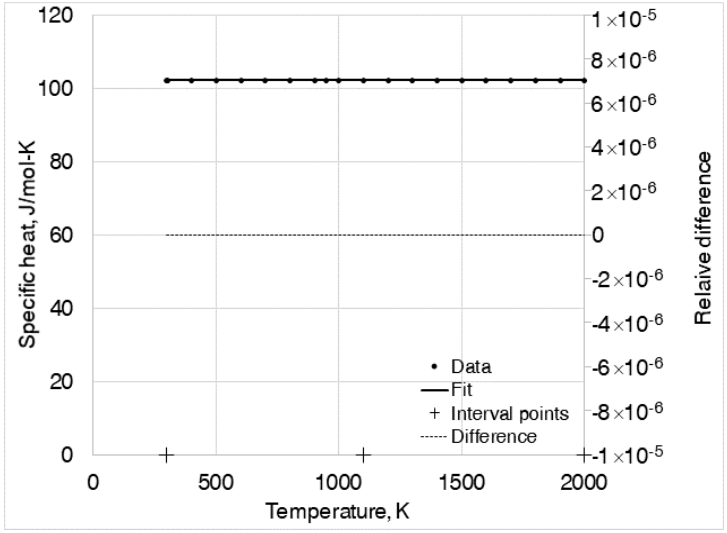


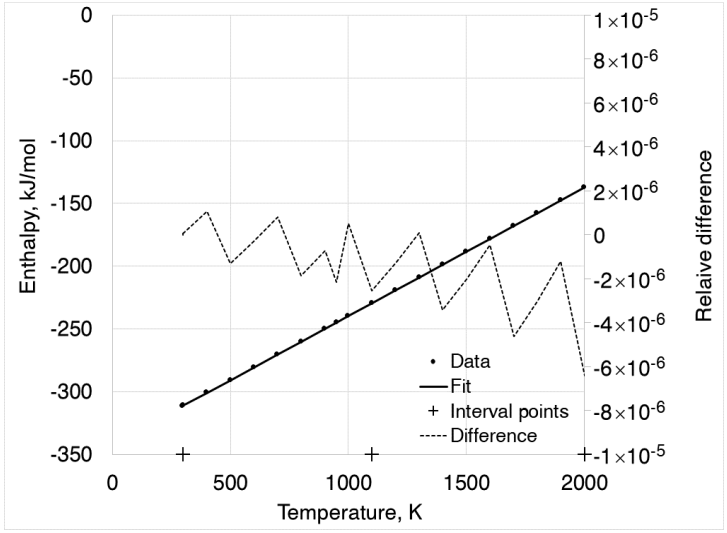


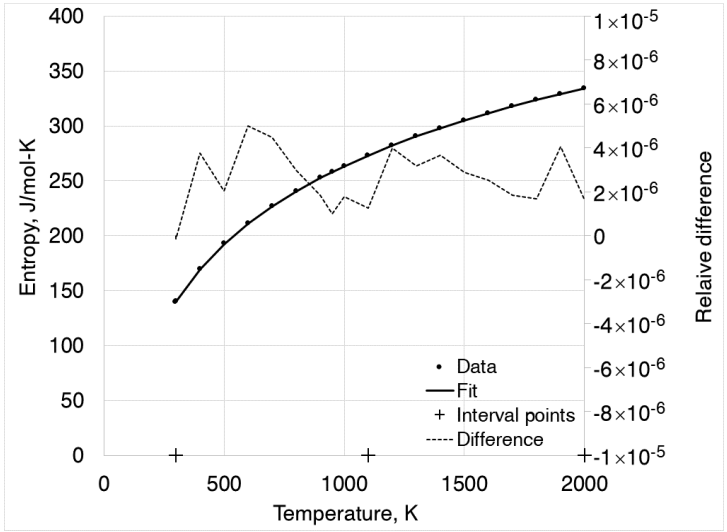


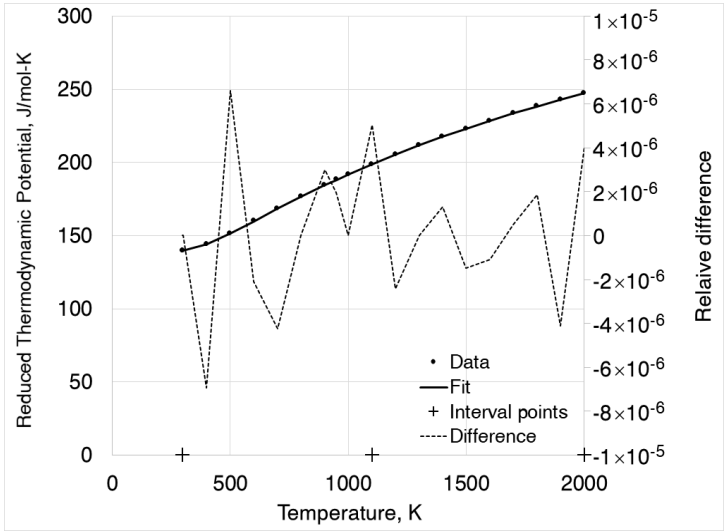
Iron Chloride ($FeCl_2$) – Cl2Fe1(l)

MW = 126.753 g/mol – $\Delta_f H^o$(298.15 K) = -311.336 kJ/mol – S^o(298.15 K) = 139.876 J/mol-K – p^o = 0.1 MPa



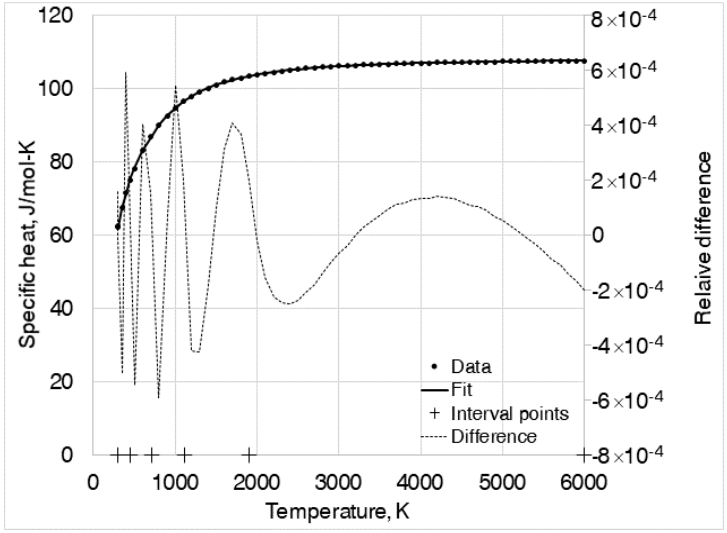


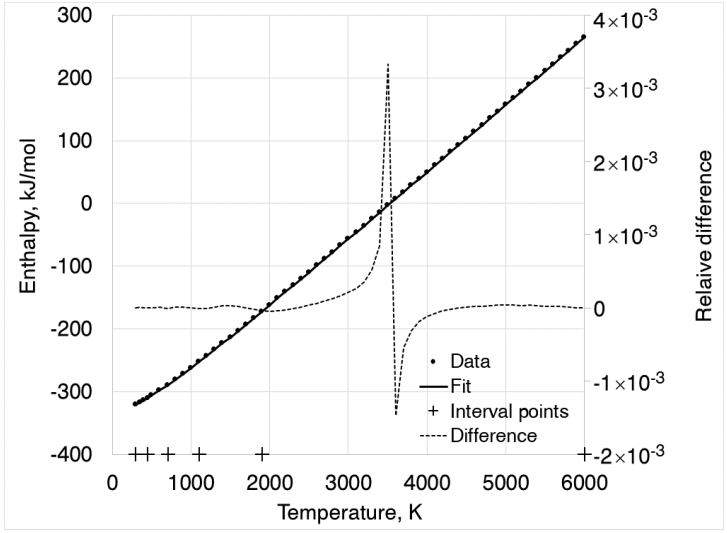


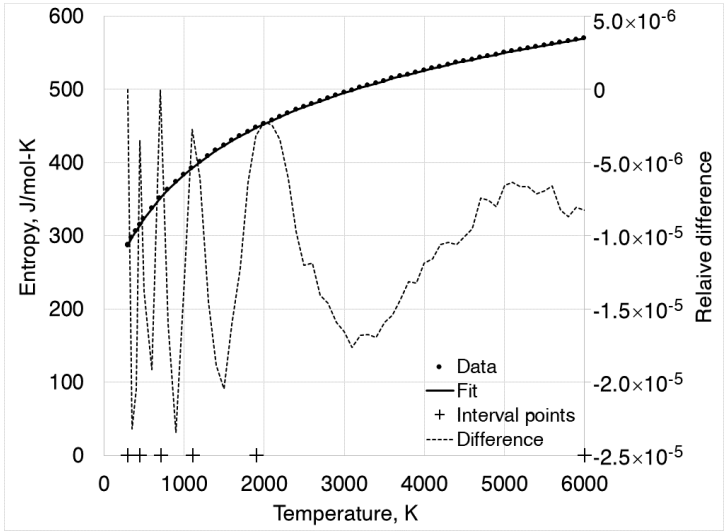


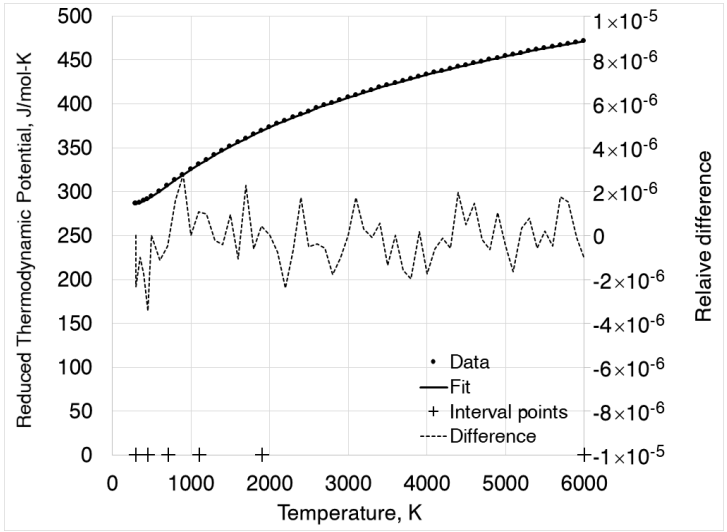
Dichlorosilane (SiH_2Cl_2) – Cl2H2Si1(g)

MW = 101.007 g/mol – $\Delta_f H^o$(298.15 K) = -320.494 kJ/mol – S^o(298.15 K) = 286.734 J/mol-K – p^o = 0.1 MPa



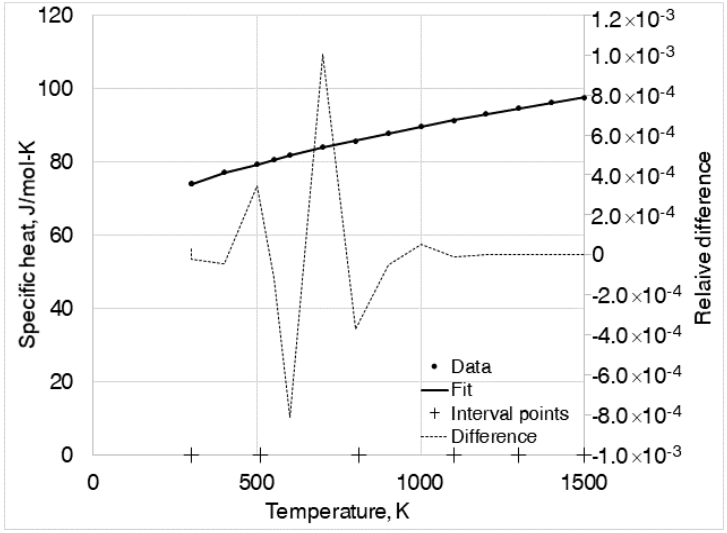


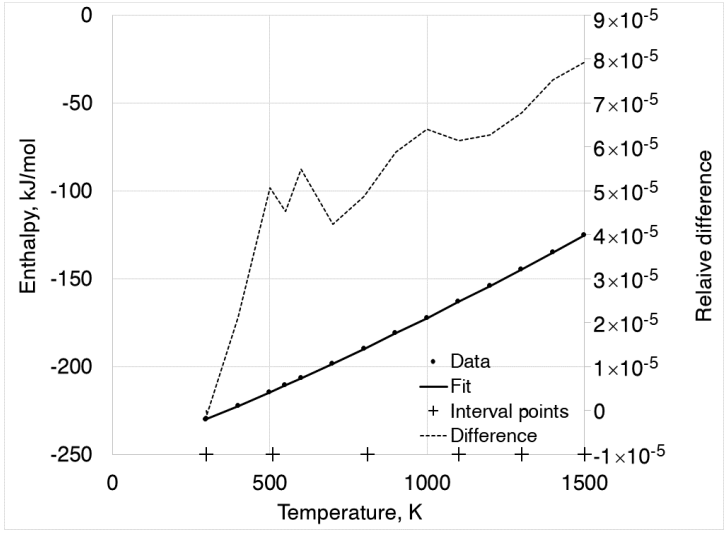


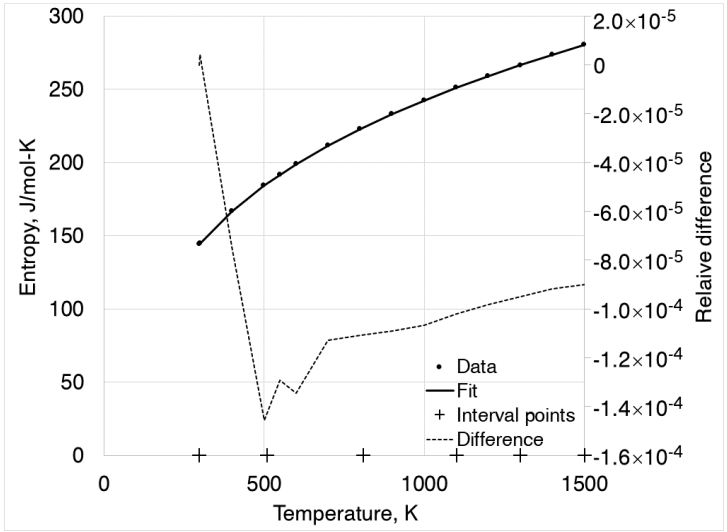


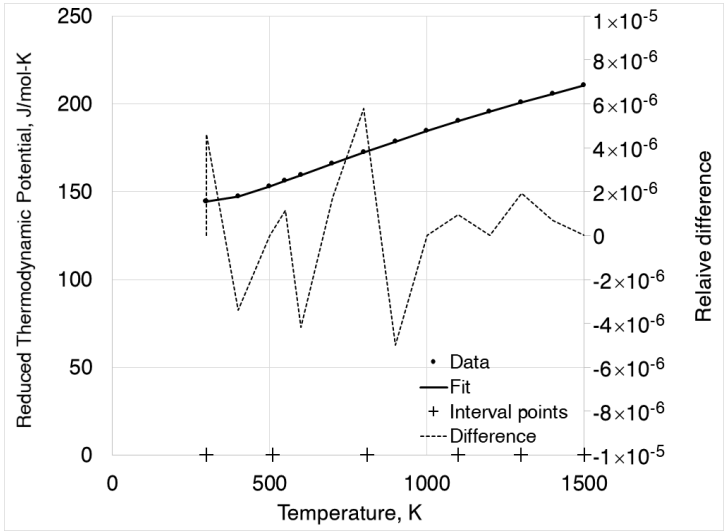
Mercury Chloride ($HgCl_2$) – Cl2Hg1(cr)

MW = 271.5 g/mol – $\Delta_f H^o$(298.15 K) = -230.12 kJ/mol – S^o(298.15 K) = 144.494 J/mol-K – p^o = 0.1 MPa



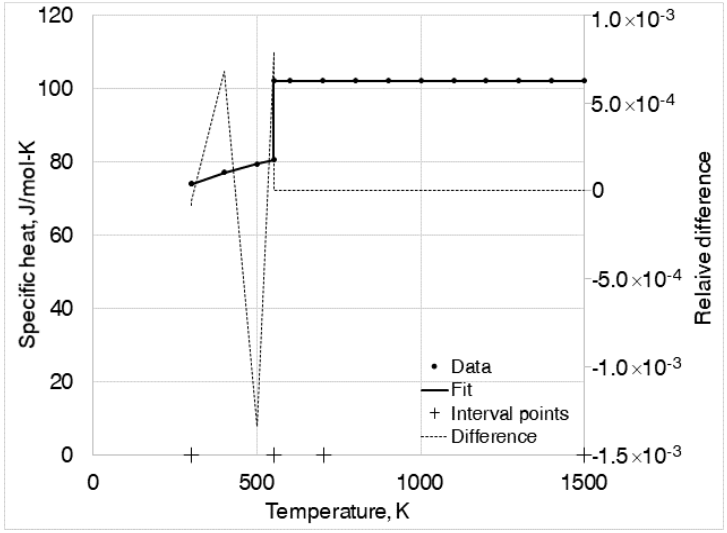


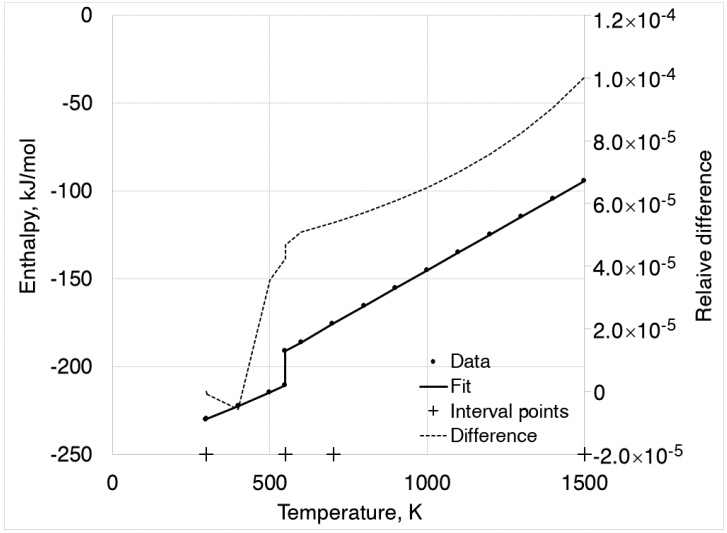


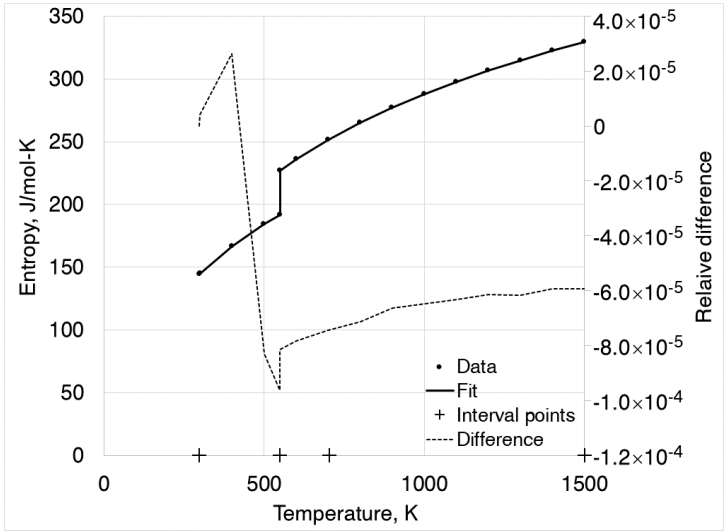


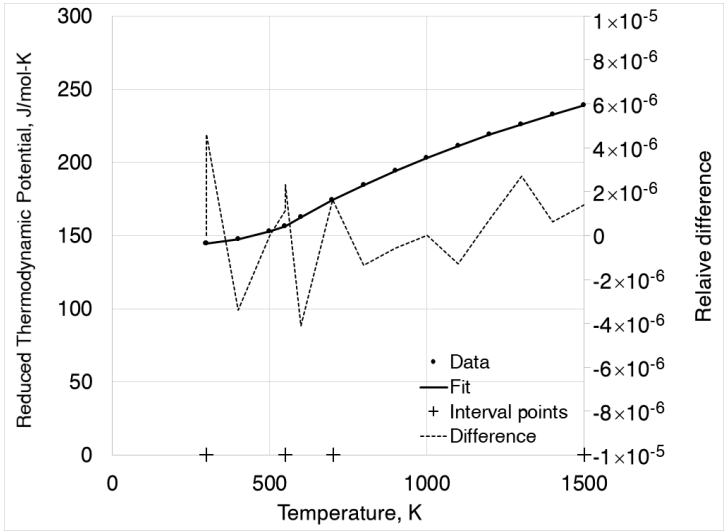
Mercury Chloride ($HgCl_2$) – Cl2Hg1(cr,l)

MW = 271.5 g/mol – $\Delta_f H^o$(298.15 K) = -230.12 kJ/mol – S^o(298.15 K) = 144.494 J/mol-K – p^o = 0.1 MPa



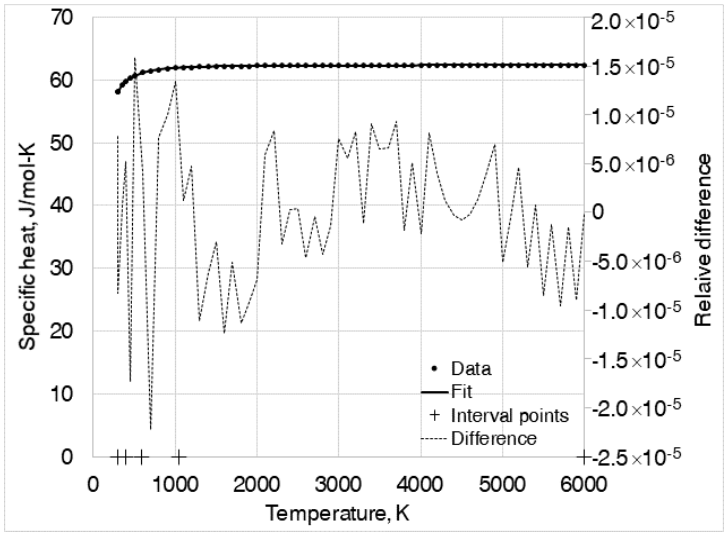


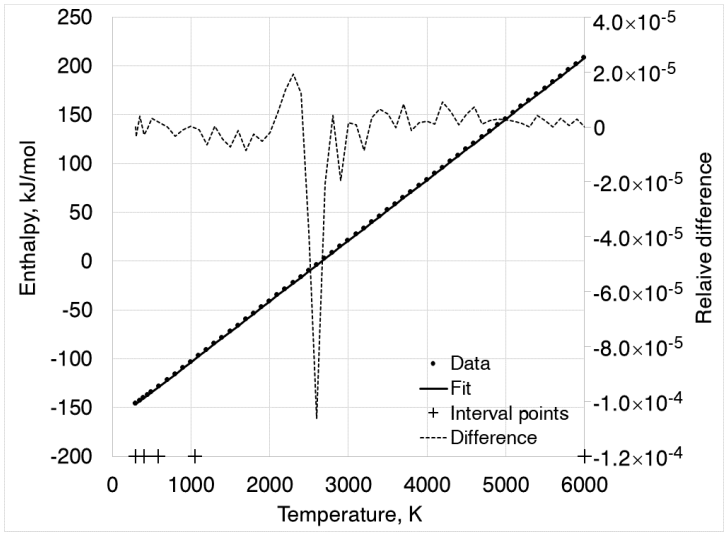


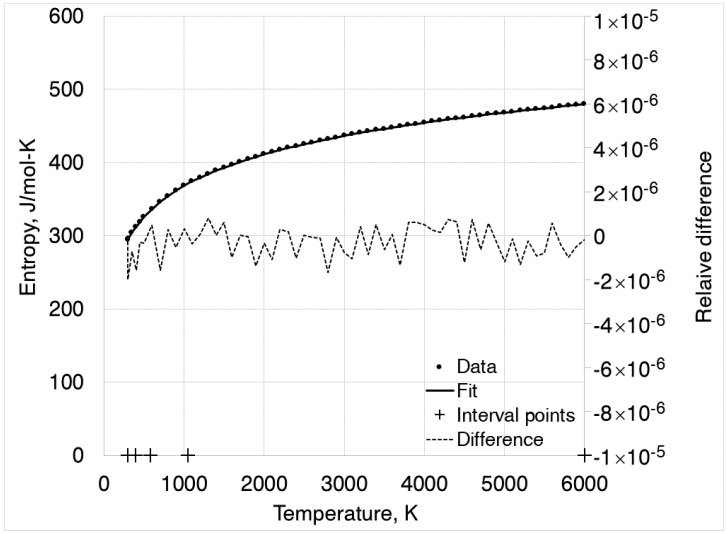


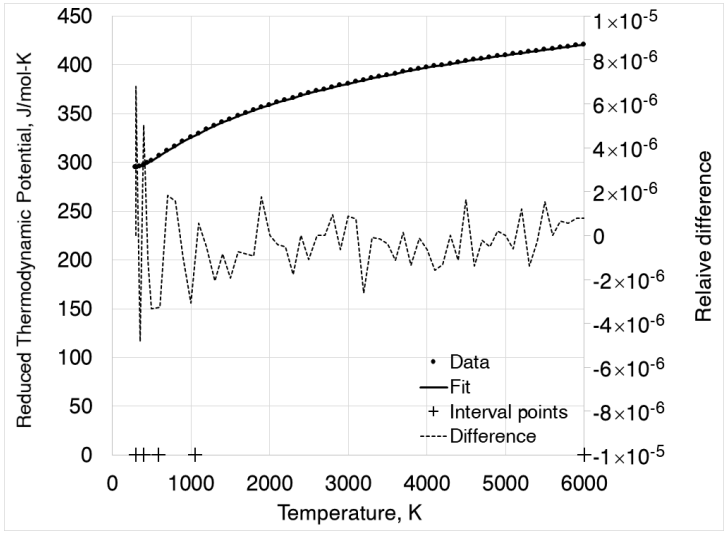
Mercury Chloride ($HgCl_2$) – Cl2Hg1(g)

MW = 271.5 g/mol – $\Delta_f H^o$(298.15 K) = -146.294 kJ/mol – S^o(298.15 K) = 294.793 J/mol-K – p^o = 0.1 MPa



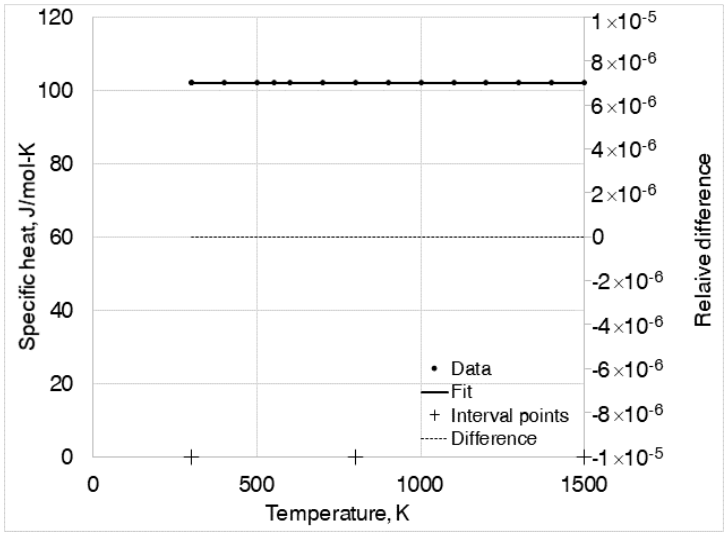


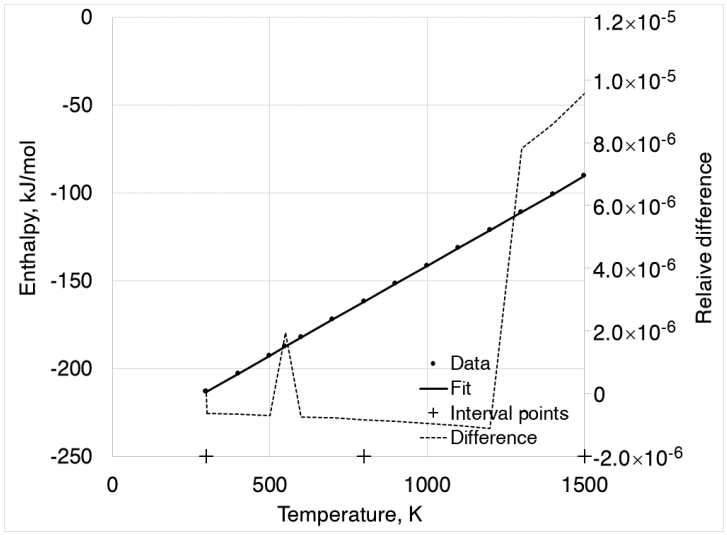


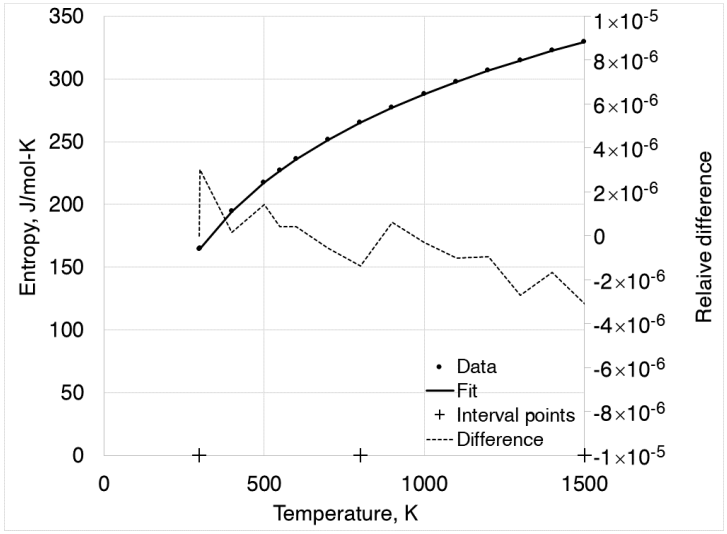


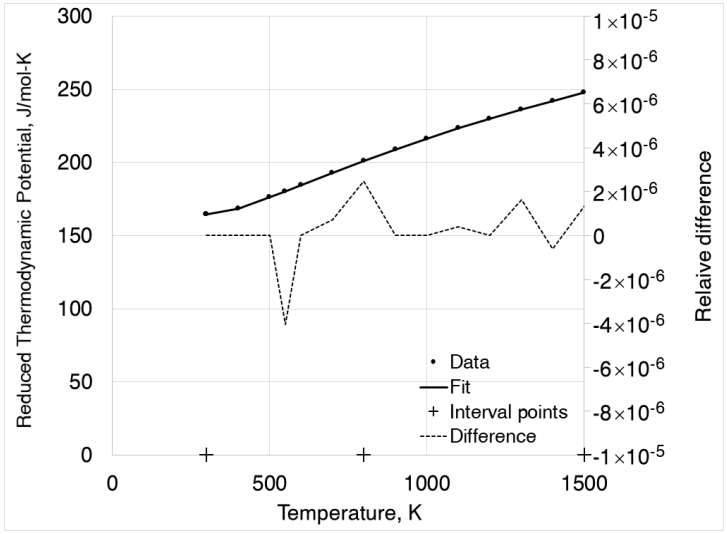
Mercury Chloride ($HgCl_2$) – Cl2Hg1(l)

MW = 271.5 g/mol – $\Delta_f H^o$(298.15 K) = -213.222 kJ/mol – S^o(298.15 K) = 164.471 J/mol-K – p^o = 0.1 MPa



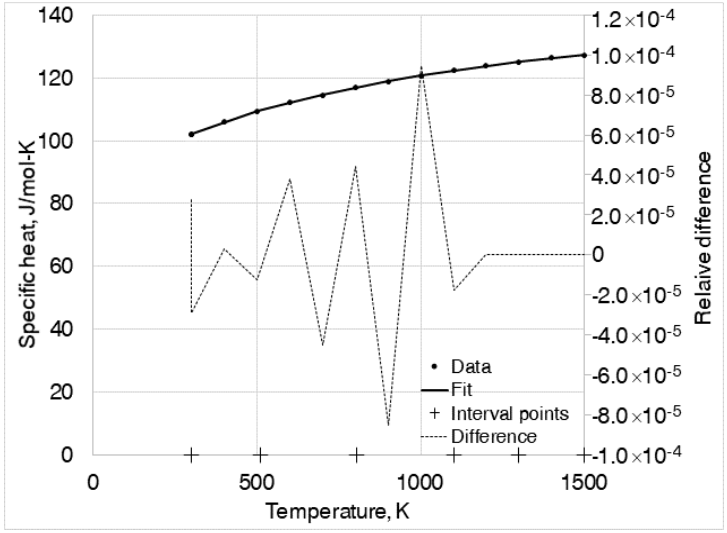


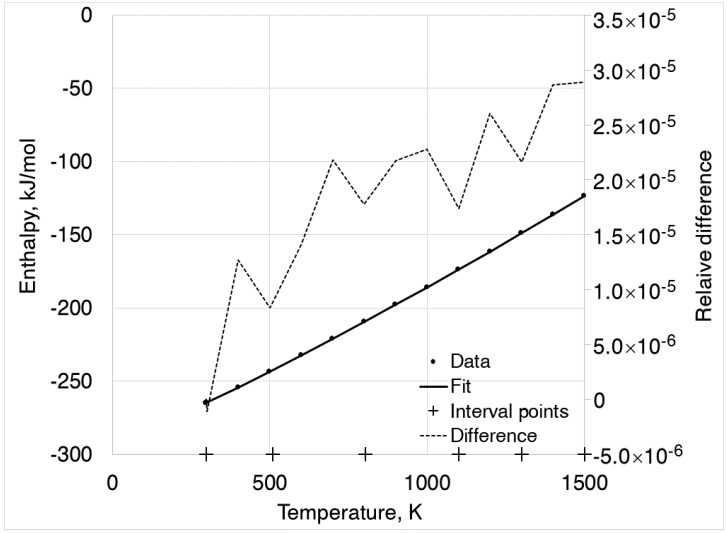


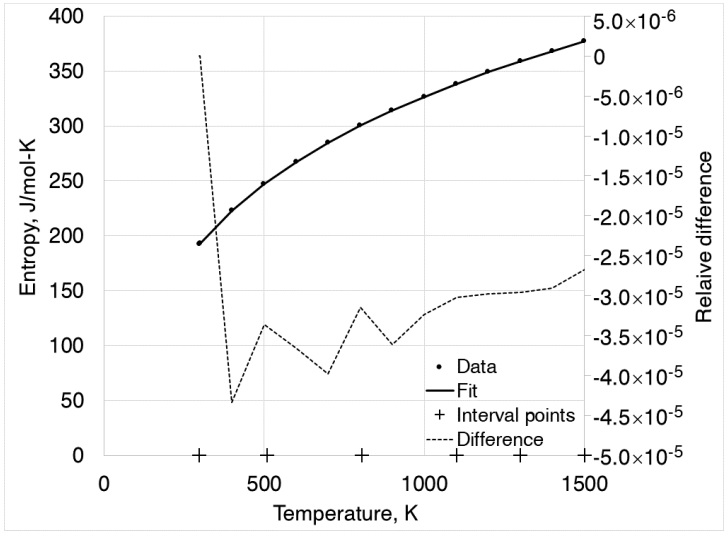


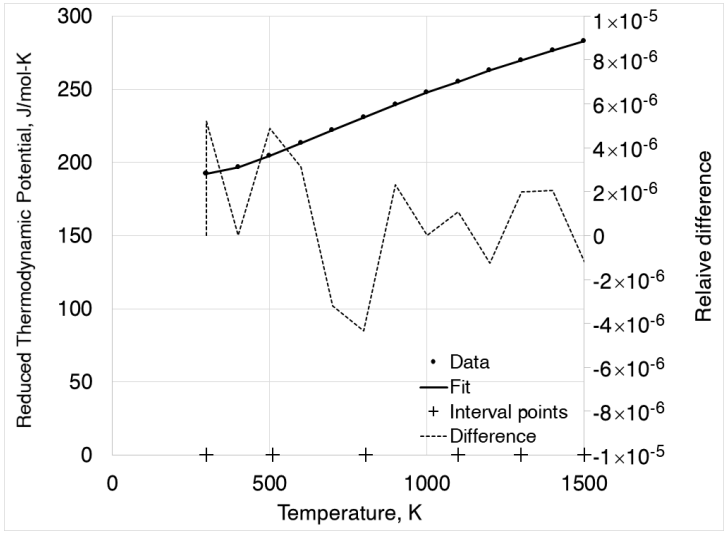
Mercury Chloride (Hg_2Cl_2) – Cl2Hg2(cr)

MW = 472.09 g/mol – $\Delta_f H^o$(298.15 K) = -264.927 kJ/mol – S^o(298.15 K) = 192.535 J/mol-K – p^o = 0.1 MPa



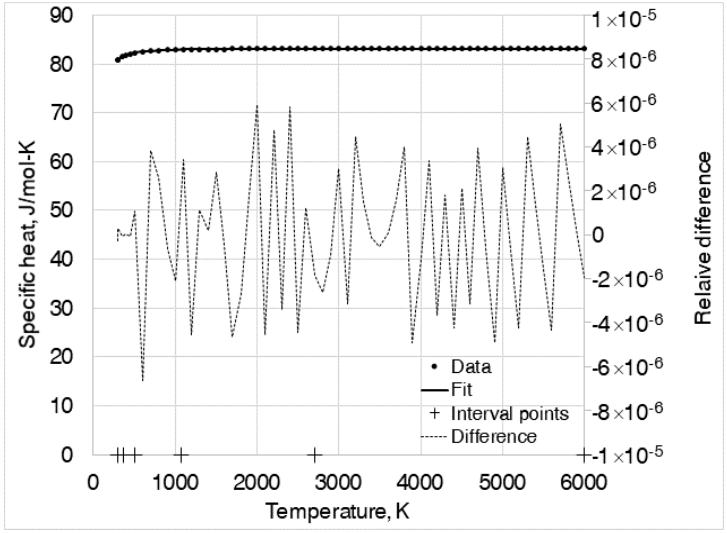


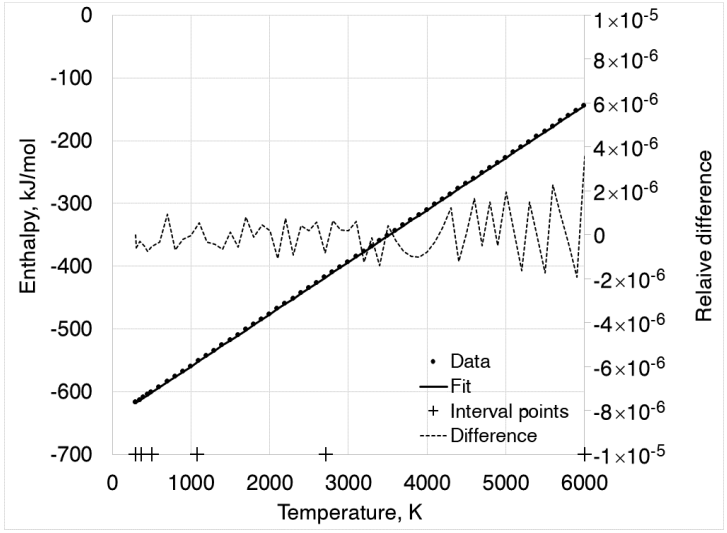


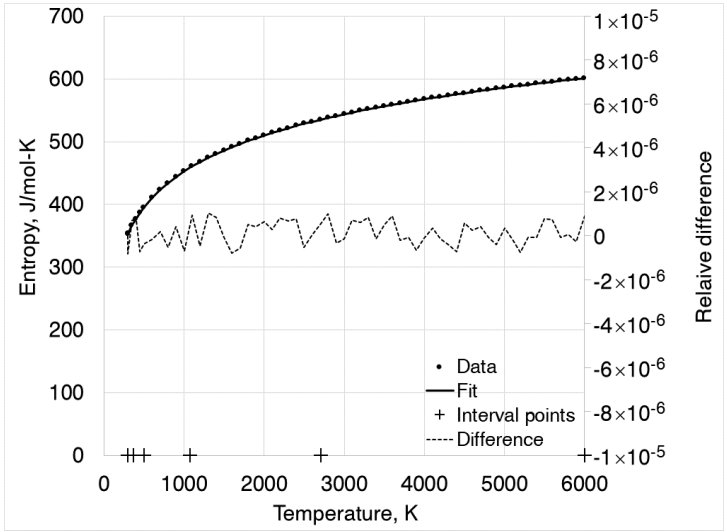


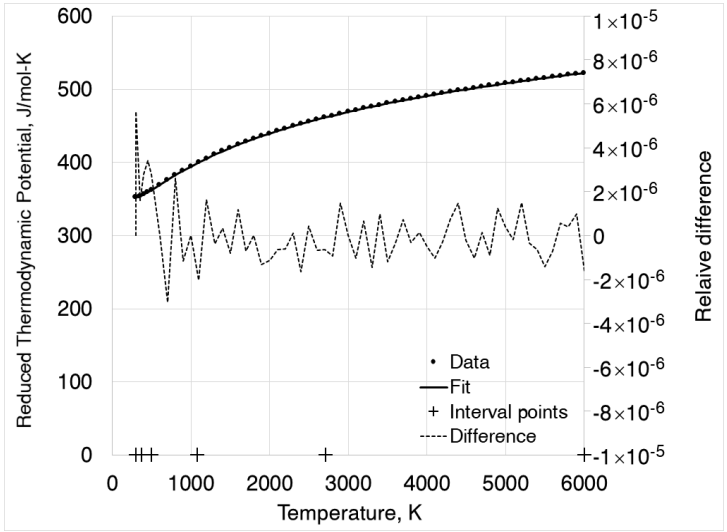
Potassium Chloride ($(KCl)_2$) – Cl2K2(g)

MW = 149.103 g/mol – $\Delta_f H^o$(298.15 K) = -617.642 kJ/mol – S^o(298.15 K) = 352.864 J/mol-K – p^o = 0.1 MPa



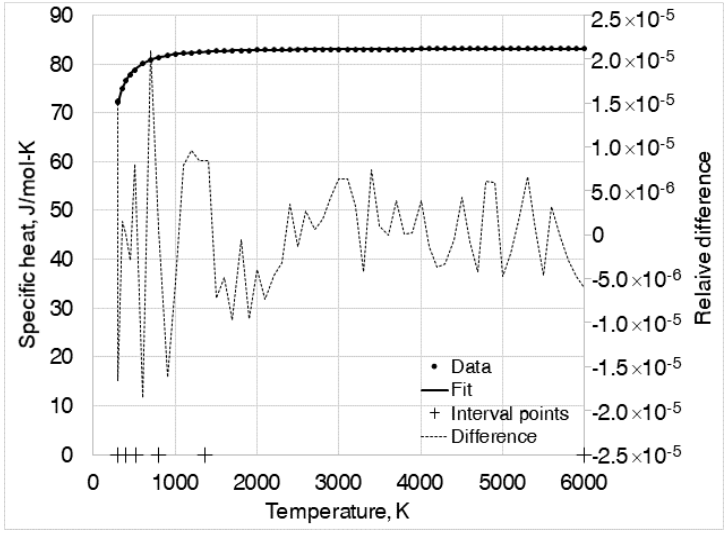


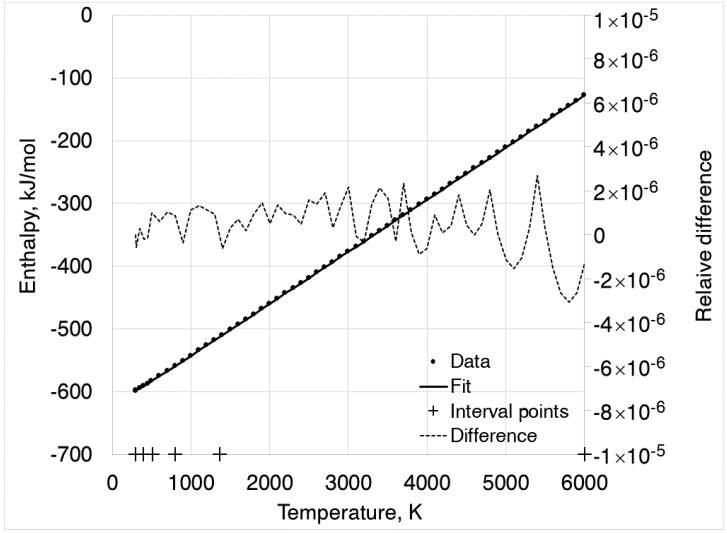


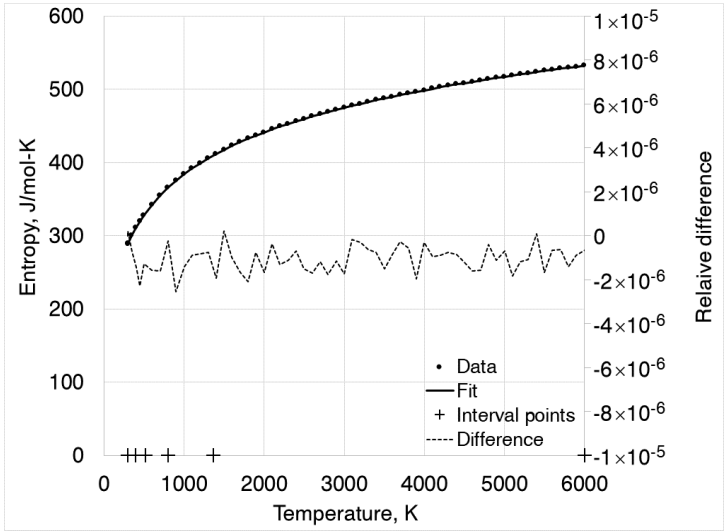


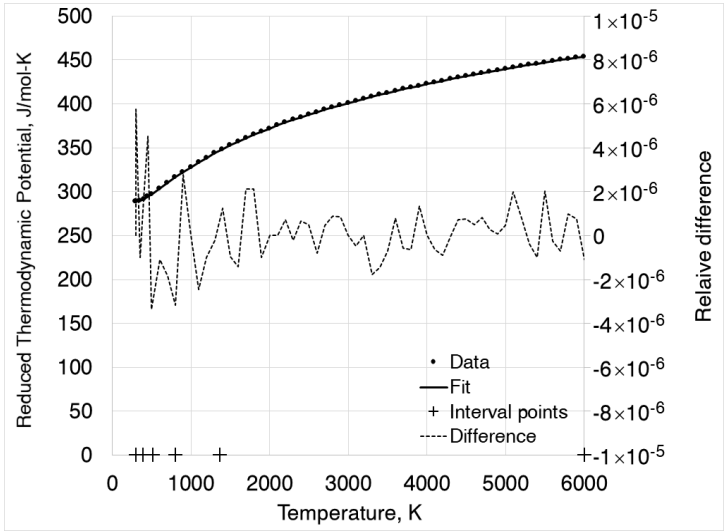
Lithium Chloride ($(LiCl)_2$) – Cl2Li2(g)

MW = 84.788 g/mol – $\Delta_f H^o$(298.15 K) = -598.546 kJ/mol – S^o(298.15 K) = 288.767 J/mol-K – p^o = 0.1 MPa



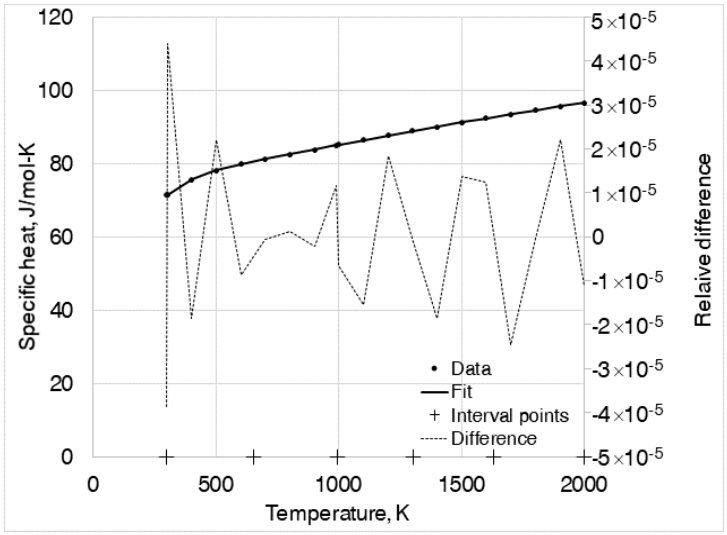


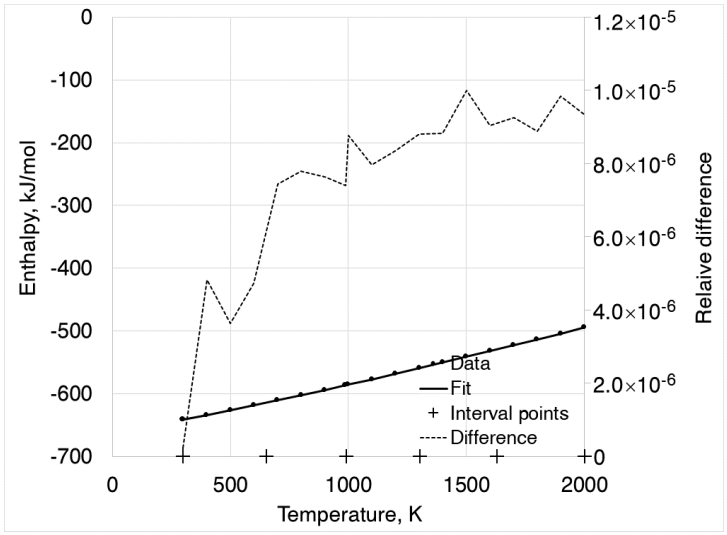


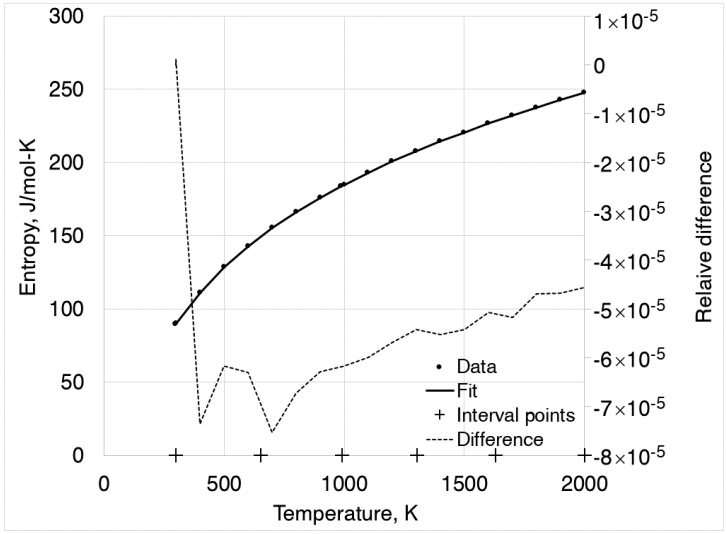


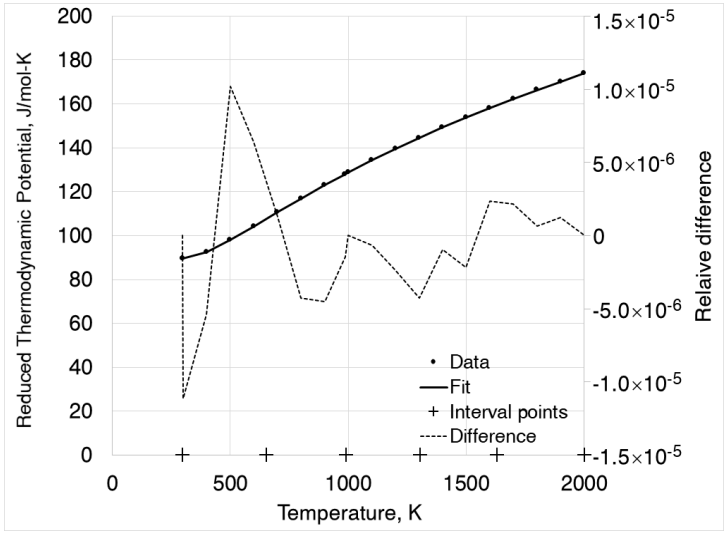
Magnesium Chloride ($MgCl_2$) – Cl2Mg1(cr)

MW = 95.211 g/mol – $\Delta_f H^o$(298.15 K) = -641.616 kJ/mol – S^o(298.15 K) = 89.629 J/mol-K – p^o = 0.1 MPa



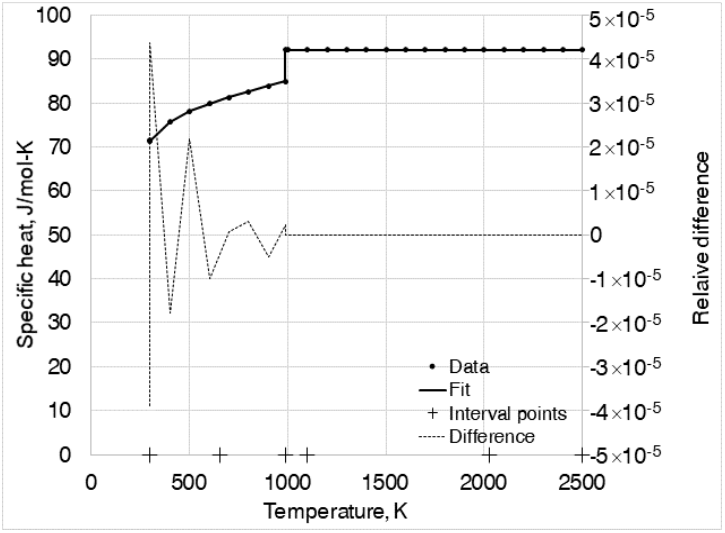


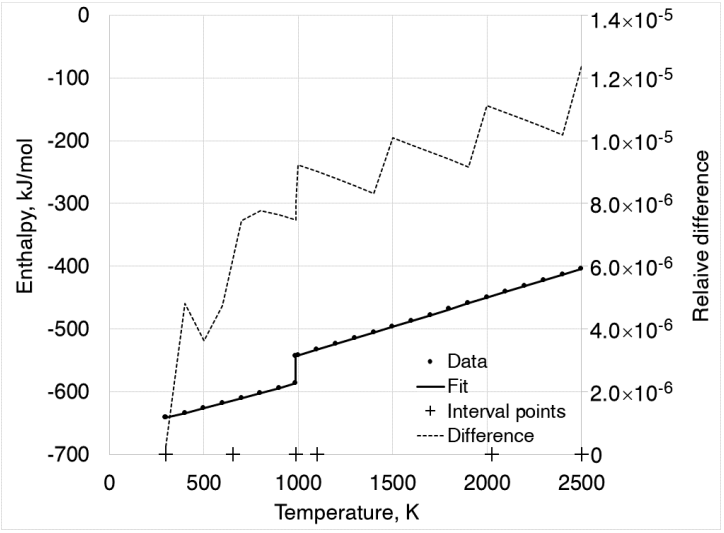


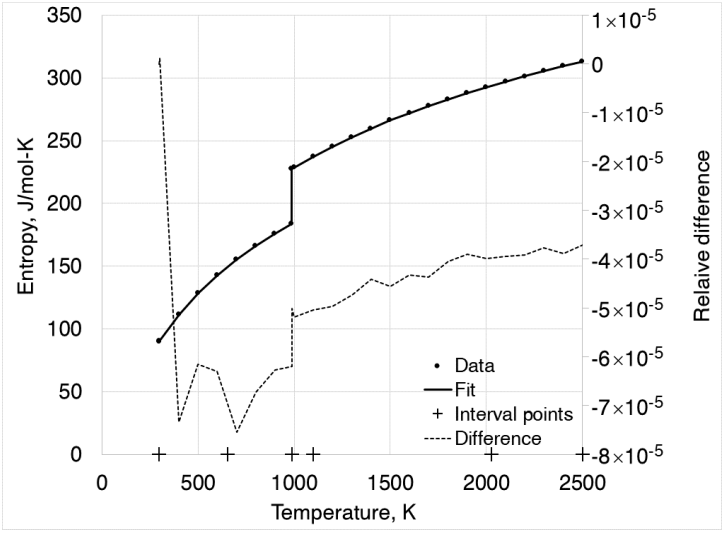


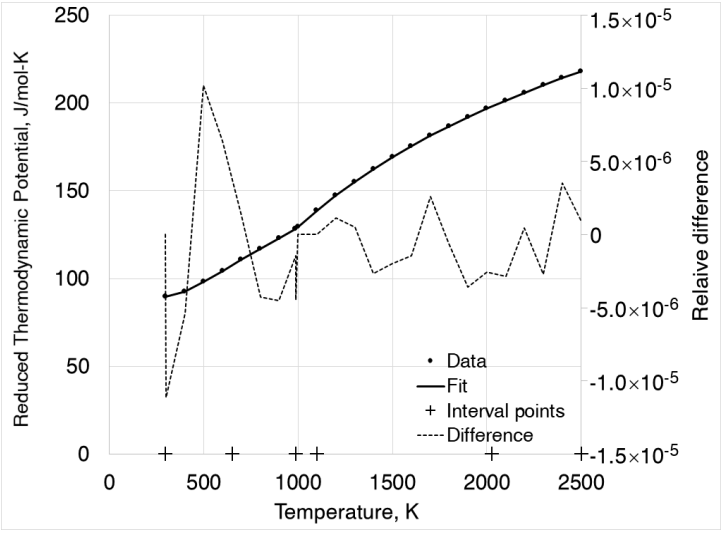
Magnesium Chloride ($MgCl_2$) – Cl2Mg1(cr,l)

MW = 95.211 g/mol – $\Delta_f H^o$(298.15 K) = -641.616 kJ/mol – S^o(298.15 K) = 89.629 J/mol-K – p^o = 0.1 MPa



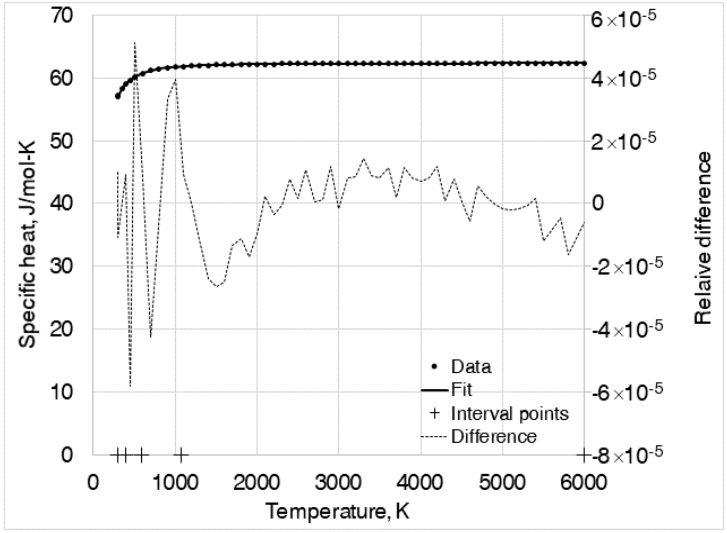


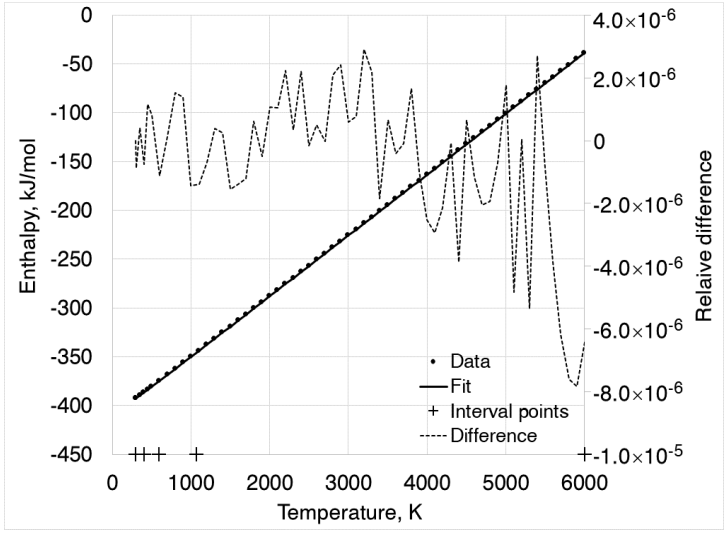


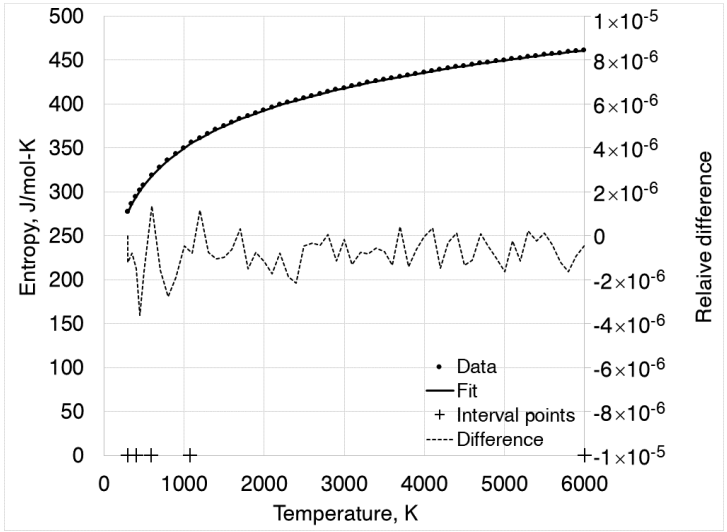


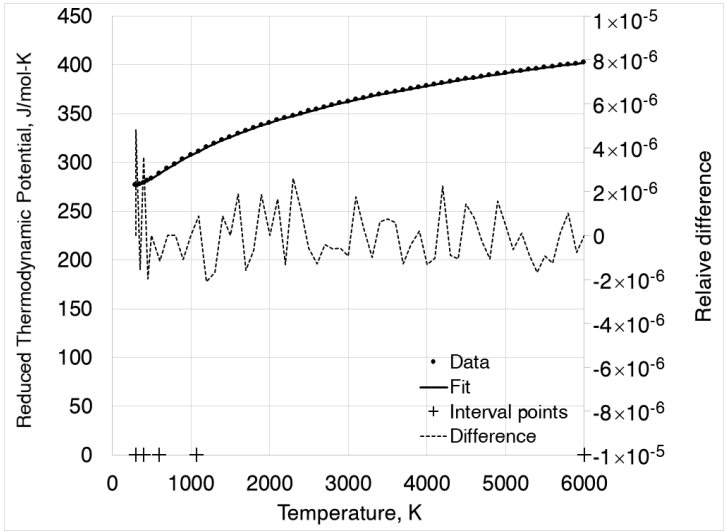
Magnesium Chloride ($MgCl_2$) – Cl2Mg1(g)

MW = 95.211 g/mol – $\Delta_f H^o$(298.15 K) = -392.459 kJ/mol – S^o(298.15 K) = 277.027 J/mol-K – p^o = 0.1 MPa



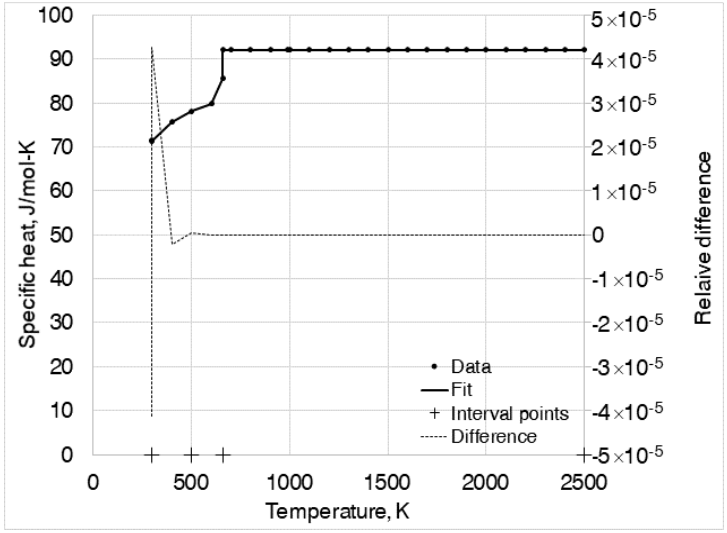


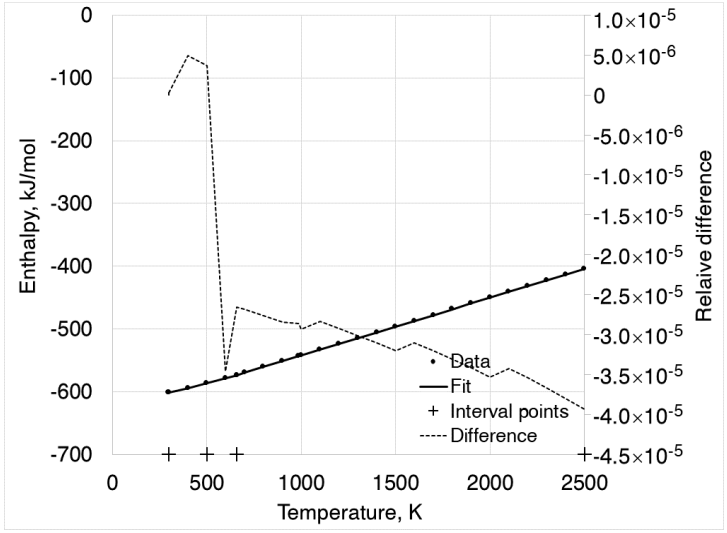


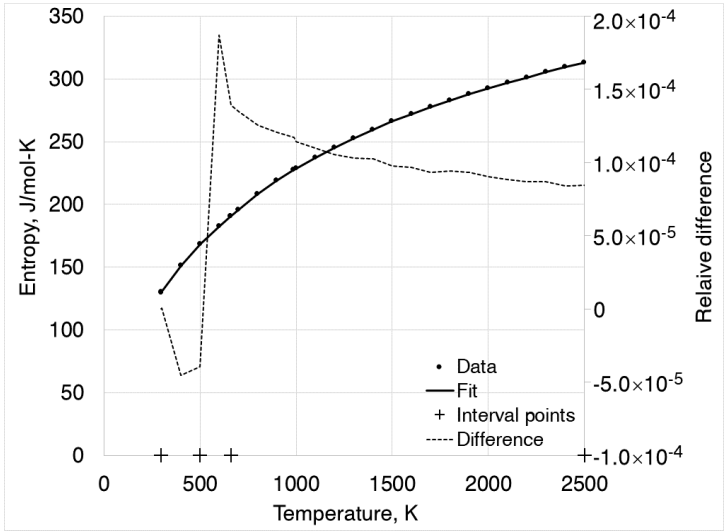


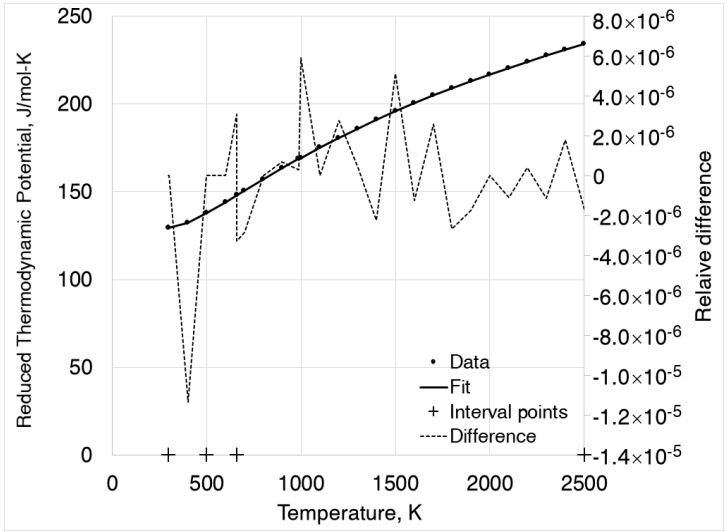
Magnesium Chloride ($MgCl_2$) – Cl2Mg1(l)

MW = 95.211 g/mol – $\Delta_f H^o$(298.15 K) = -601.577 kJ/mol – S^o(298.15 K) = 129.48 J/mol-K – p^o = 0.1 MPa



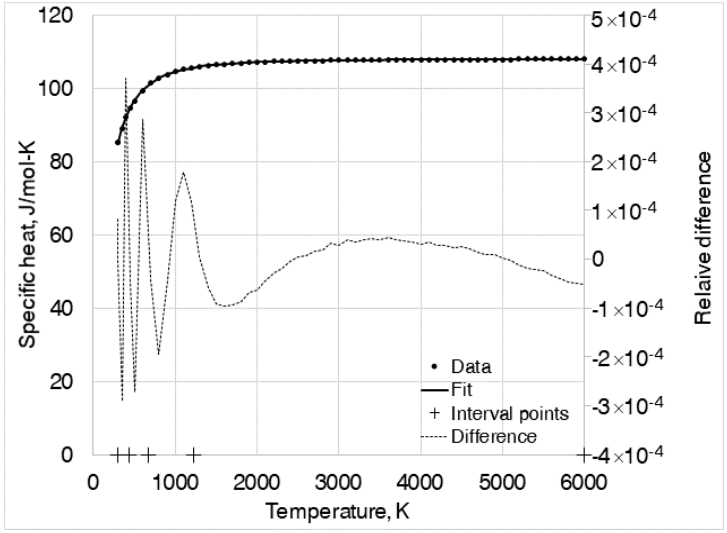


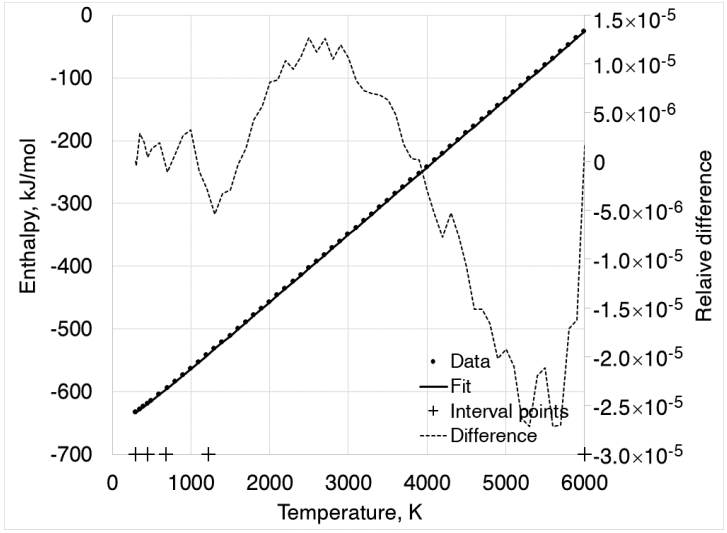


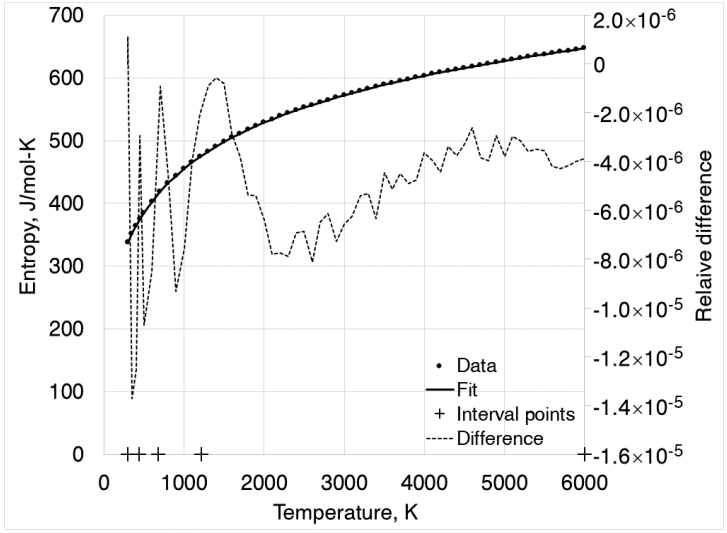


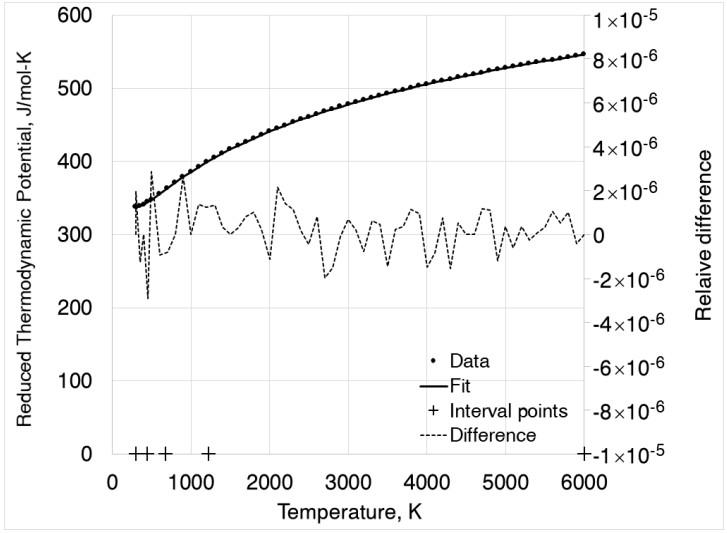
Molybdenum Chloride Oxide (MoO_2Cl_2) – Cl2Mo1O2(g)

MW = 198.84 g/mol – $\Delta_f H^o$(298.15 K) = -633.039 kJ/mol – S^o(298.15 K) = 337.787 J/mol-K – p^o = 0.1 MPa



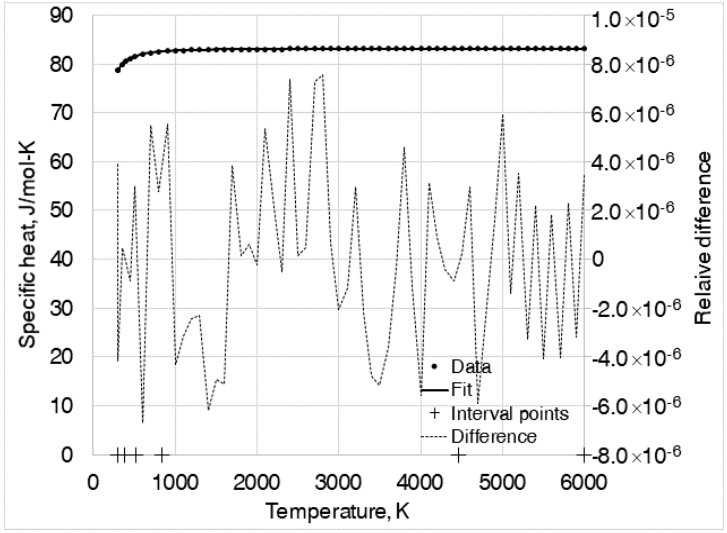


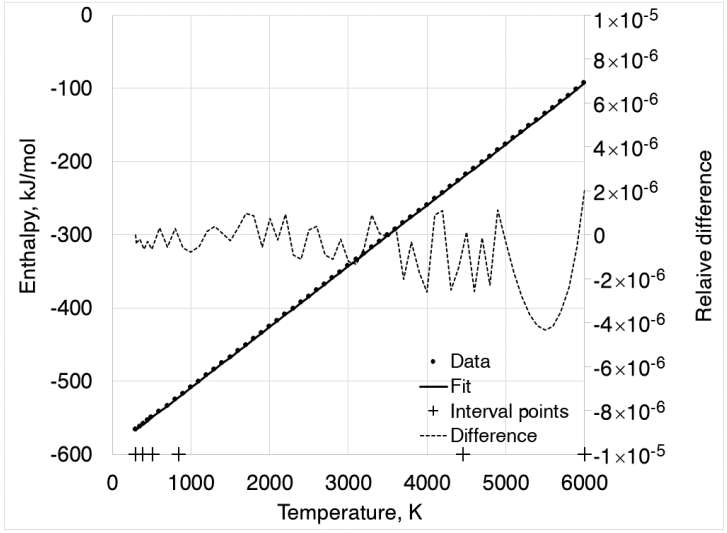


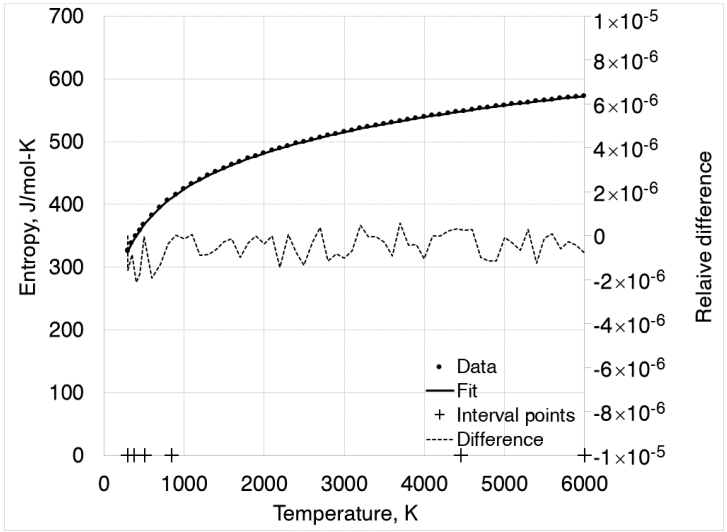


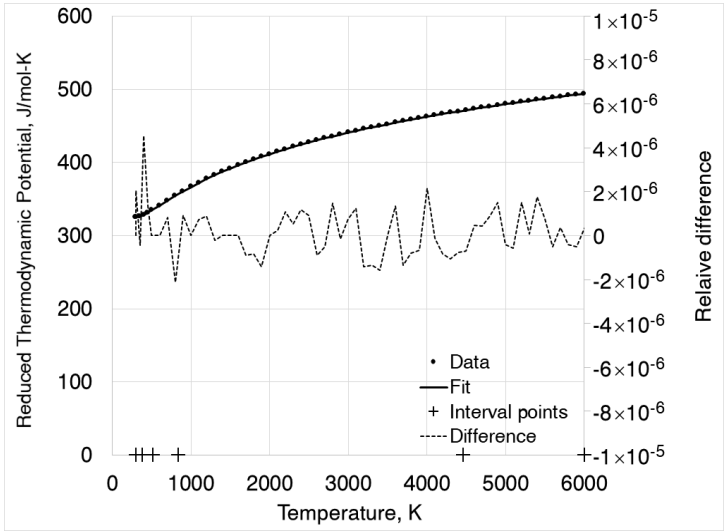
Sodium Chloride ($(NaCl)_2$) – Cl2Na2(g)

MW = 116.886 g/mol – $\Delta_f H^o$(298.15 K) = -566.095 kJ/mol – S^o(298.15 K) = 325.432 J/mol-K – p^o = 0.1 MPa



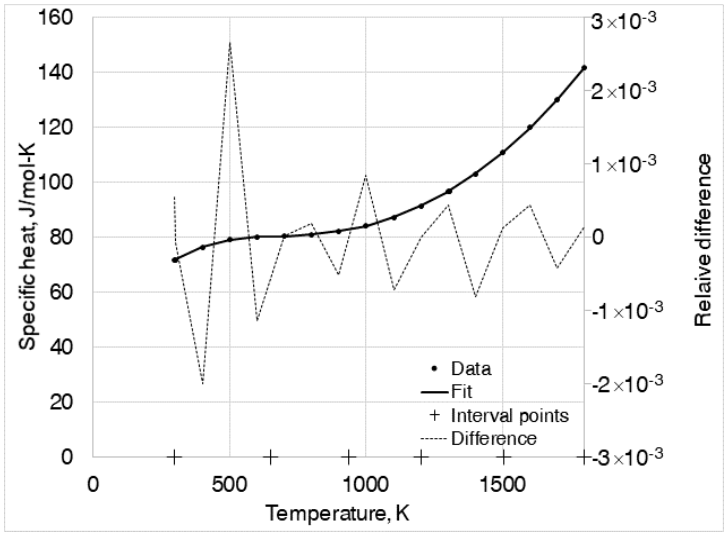


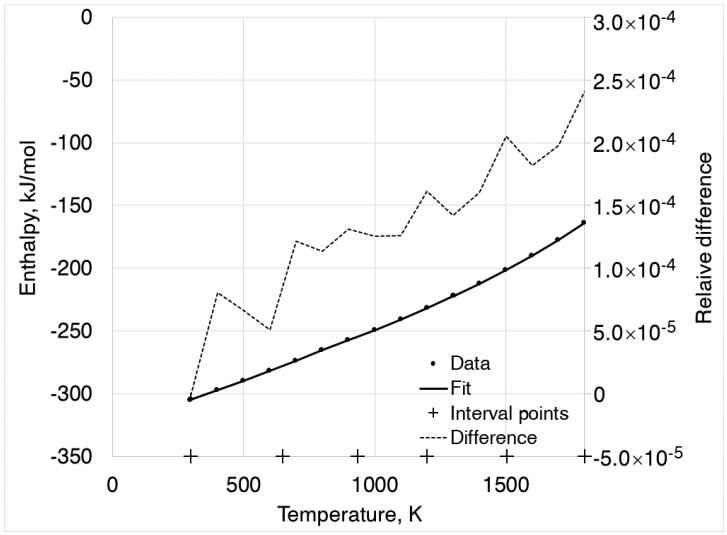


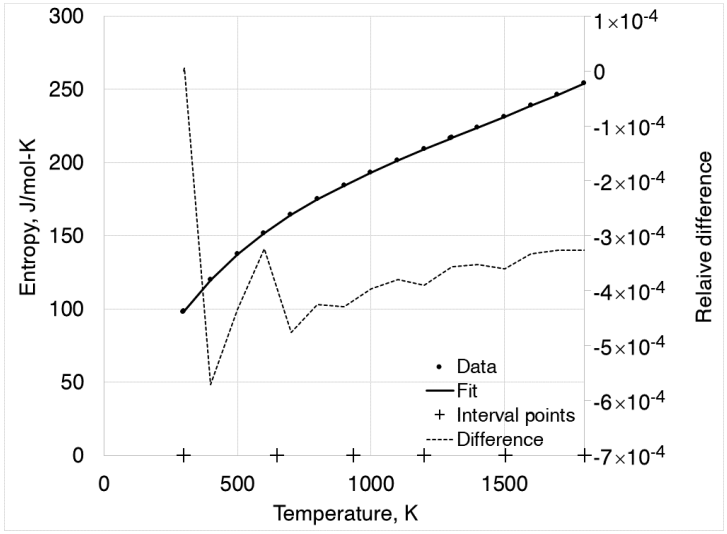


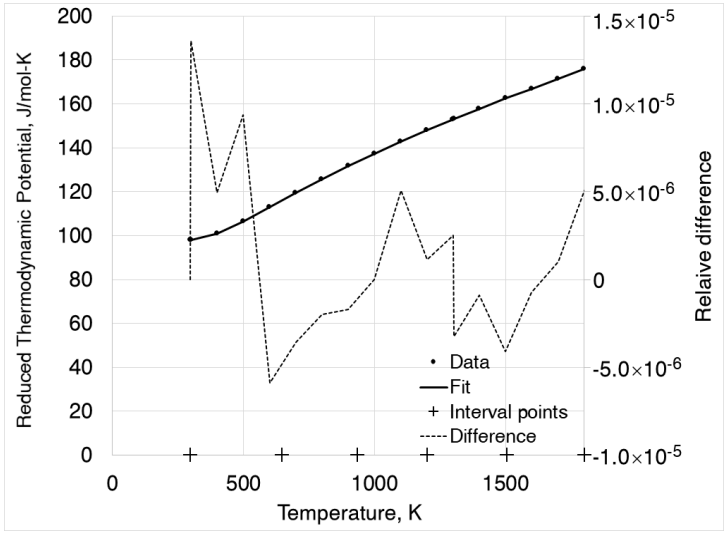
Nickel Chloride ($NiCl_2$) – Cl2Ni1(cr)

MW = 129.6 g/mol – $\Delta_f H^o$(298.15 K) = -304.93 kJ/mol – S^o(298.15 K) = 98.157 J/mol-K – p^o = 0.1 MPa



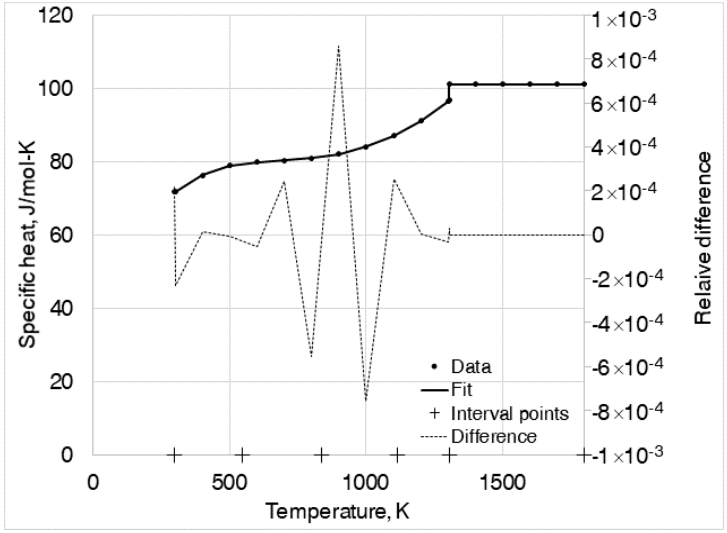


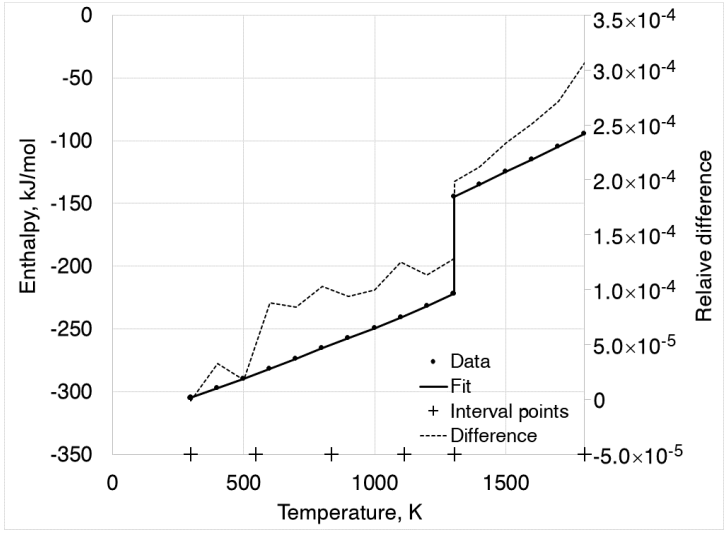


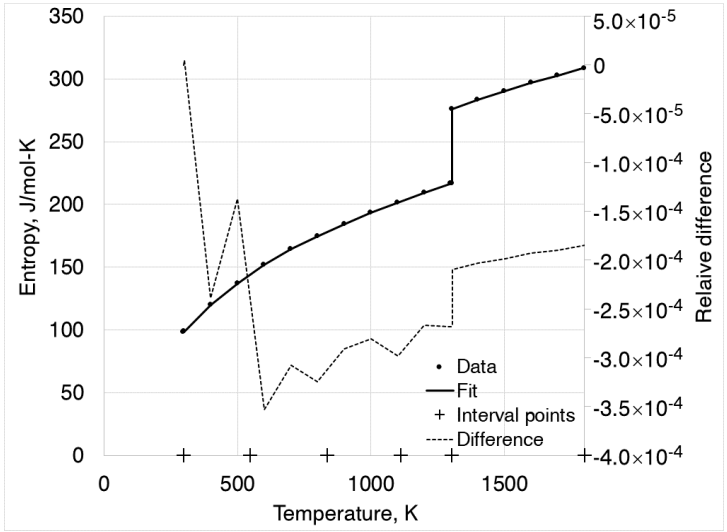


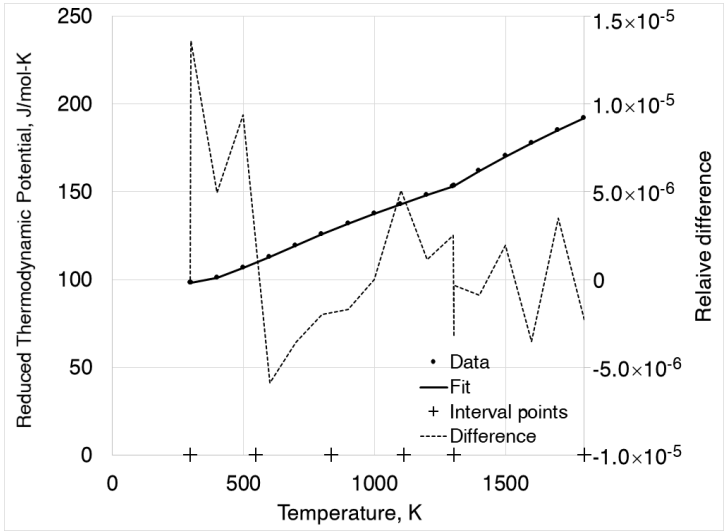
Nickel Chloride ($NiCl_2$) – Cl2Ni1(cr,l)

MW = 129.6 g/mol – $\Delta_f H^o$(298.15 K) = -304.93 kJ/mol – S^o(298.15 K) = 98.157 J/mol-K – p^o = 0.1 MPa



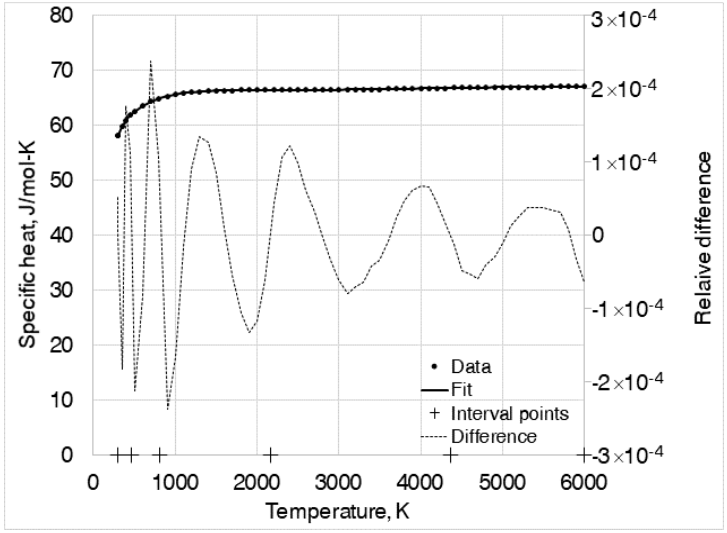


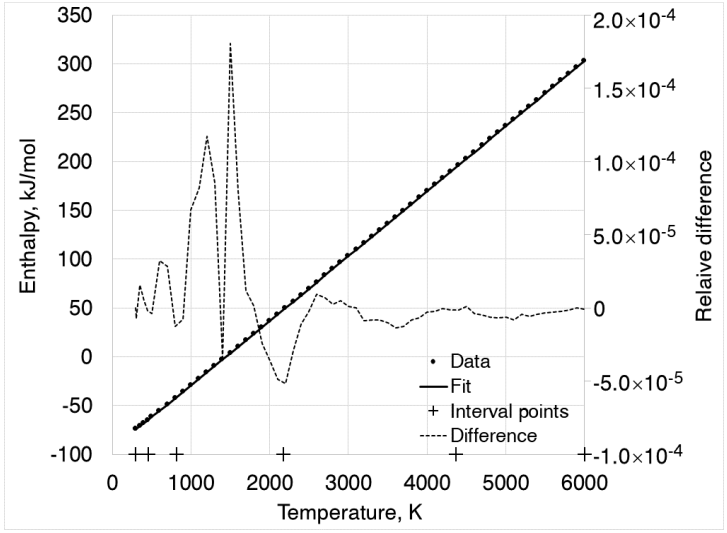


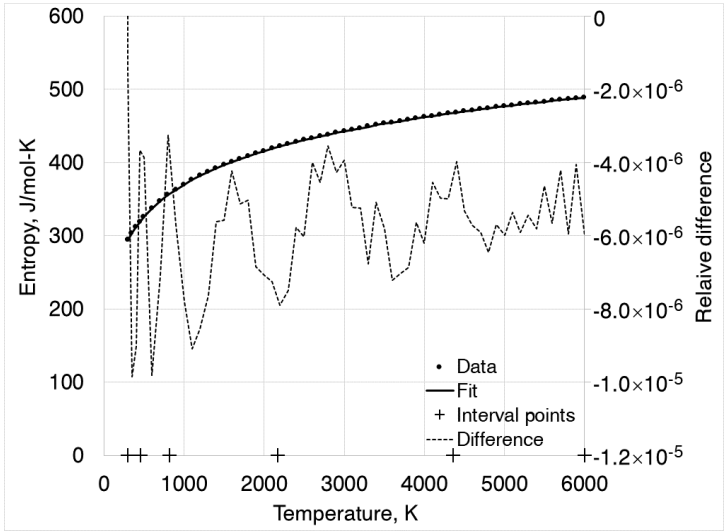


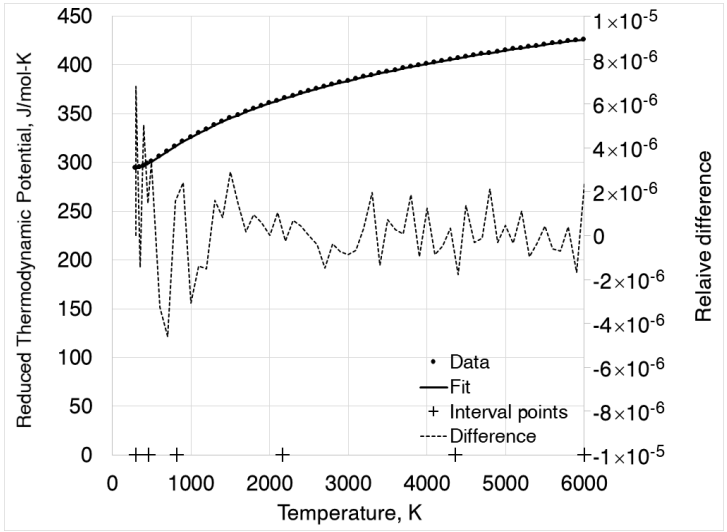
Nickel Chloride ($NiCl_2$) – Cl2Ni1(g)

MW = 129.6 g/mol – $\Delta_f H^o$(298.15 K) = -73.931 kJ/mol – S^o(298.15 K) = 294.233 J/mol-K – p^o = 0.1 MPa



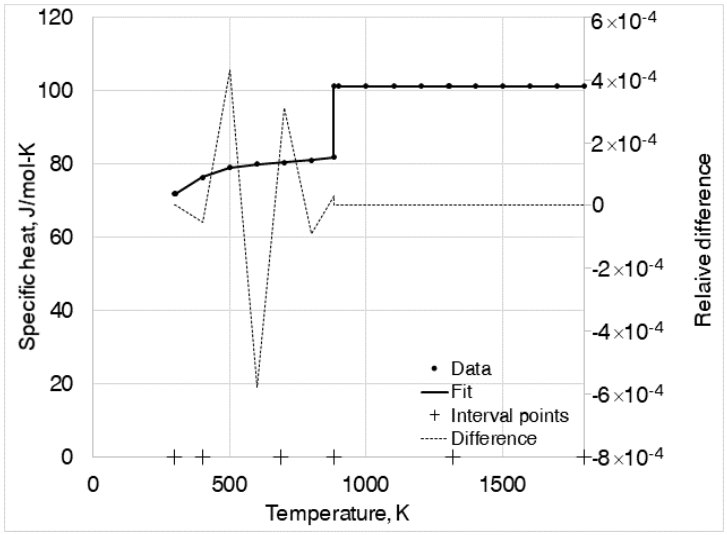


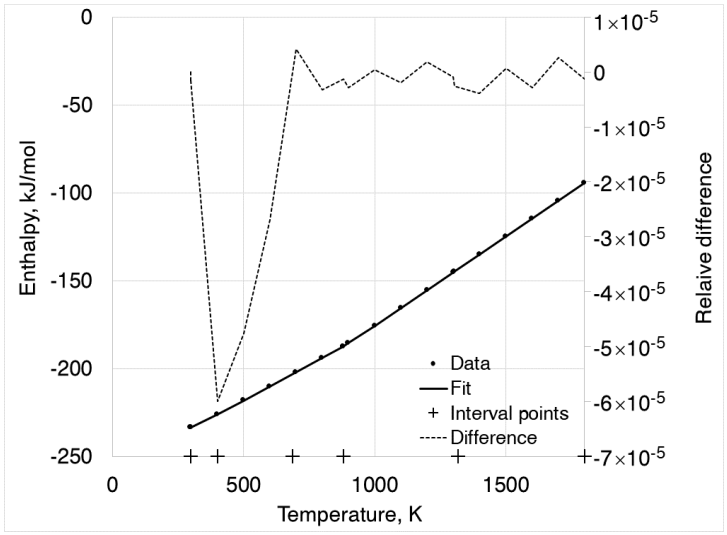


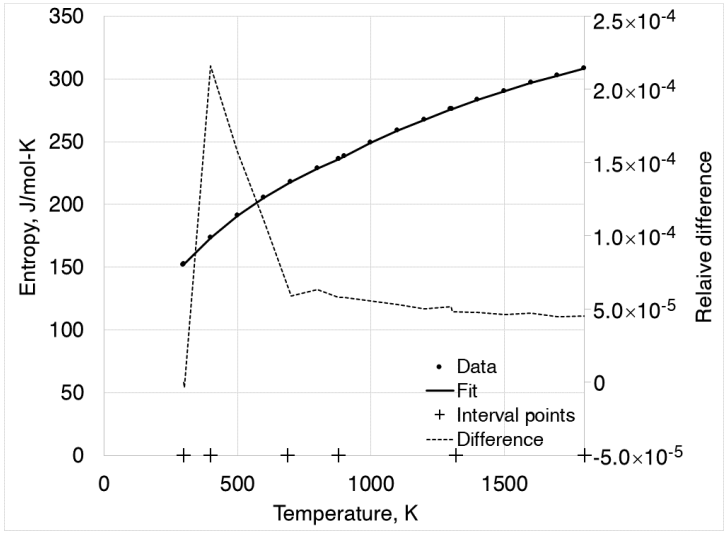


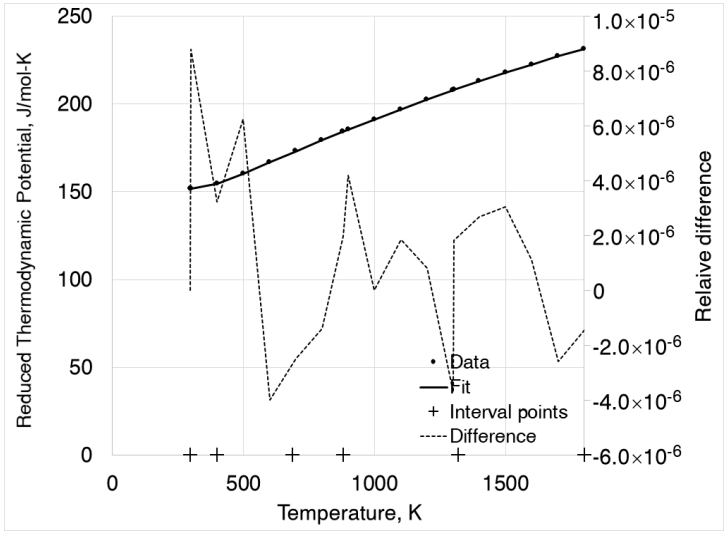
Nickel Chloride ($NiCl_2$) – Cl2Ni1(l)

MW = 129.6 g/mol – $\Delta_f H^o$(298.15 K) = -233.517 kJ/mol – S^o(298.15 K) = 151.81 J/mol-K – p^o = 0.1 MPa



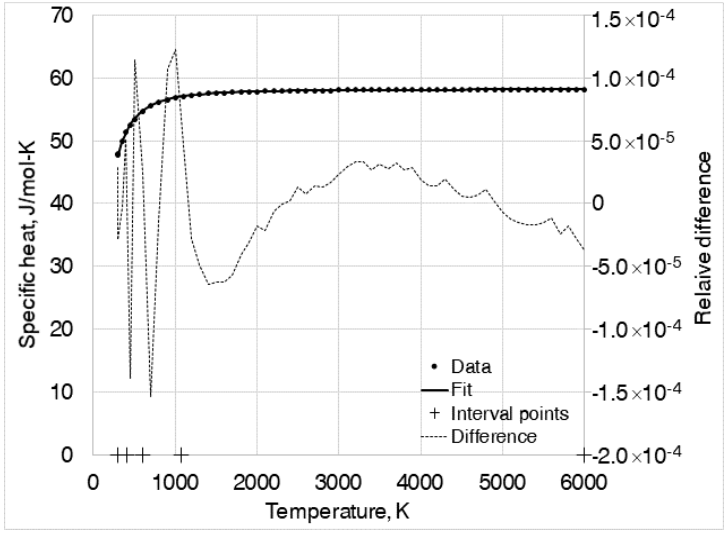


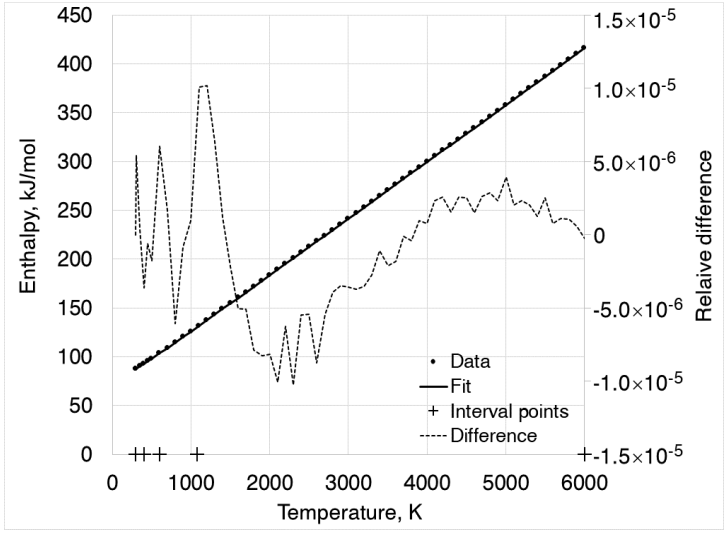


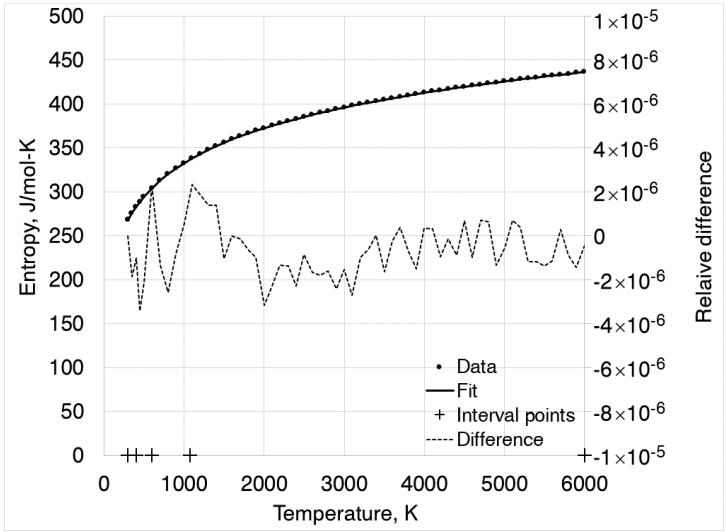


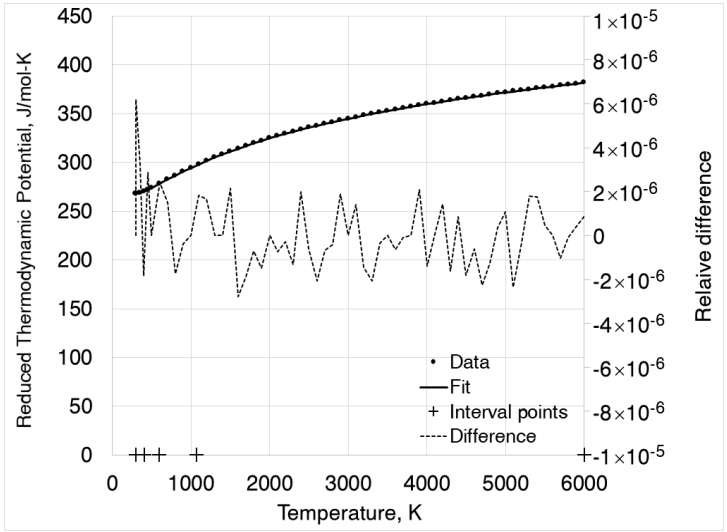
Chlorine Oxide (Cl_2O) – Cl2O1(g)

MW = 86.905 g/mol – $\Delta_f H^o$(298.15 K) = 87.864 kJ/mol – S^o(298.15 K) = 267.959 J/mol-K – p^o = 0.1 MPa



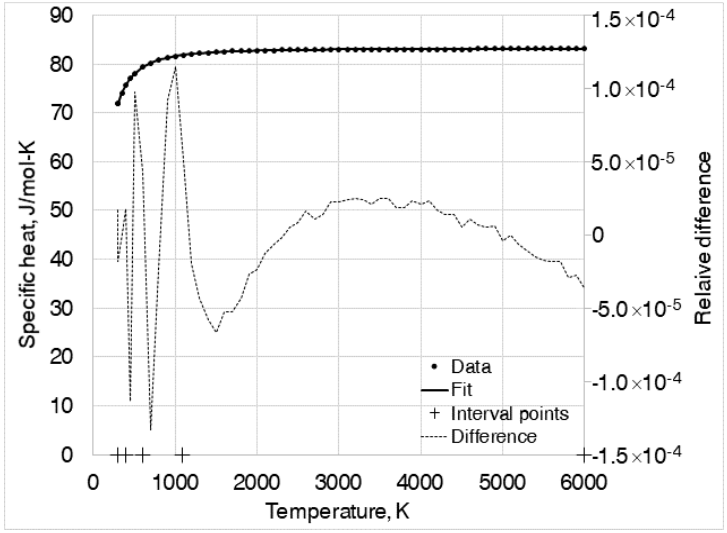


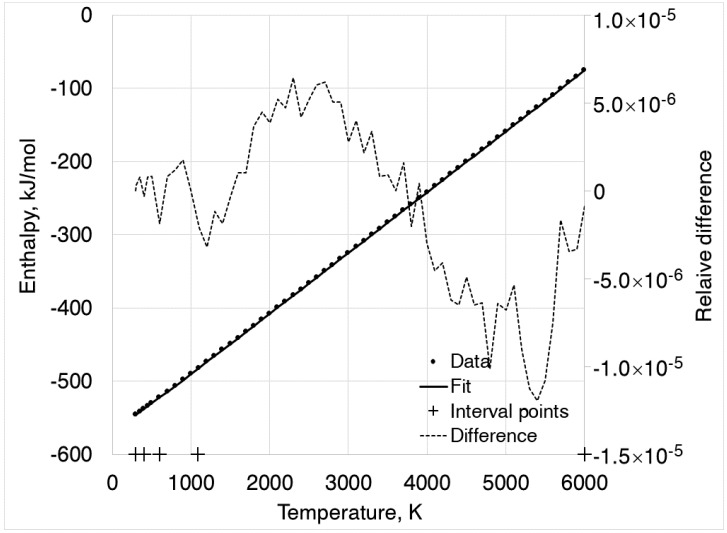


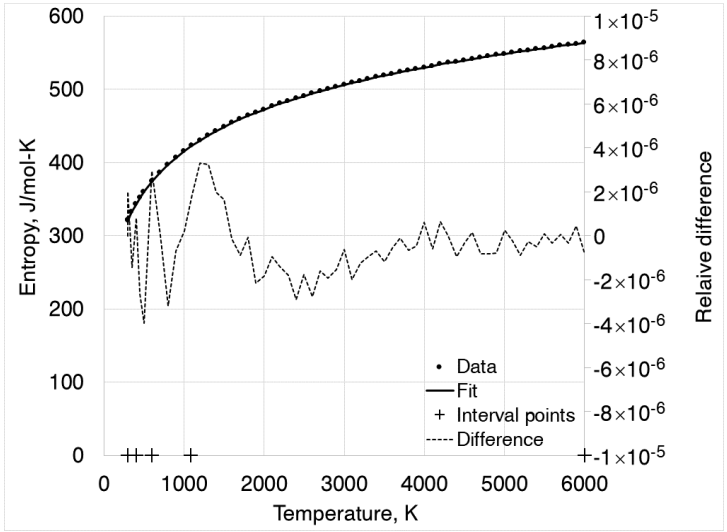


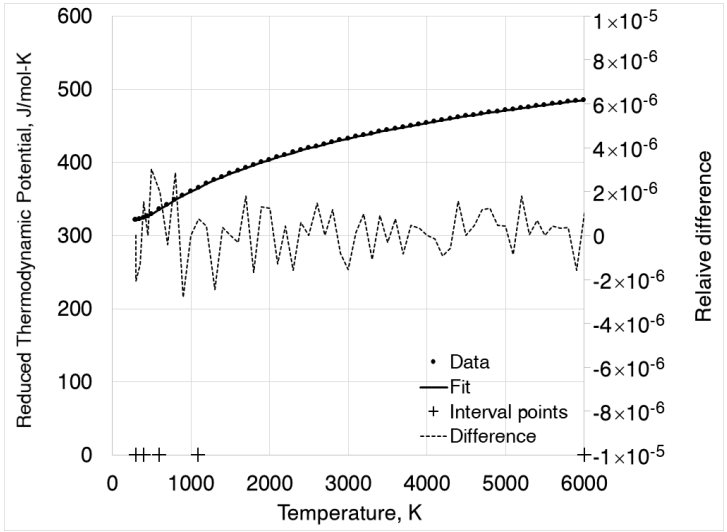
Titanium Chloride Oxide ($TiOCl_2$) – Cl2O1Ti1(g)

MW = 134.79 g/mol – $\Delta_f H^o$(298.15 K) = -545.552 kJ/mol – S^o(298.15 K) = 321.001 J/mol-K – p^o = 0.1 MPa



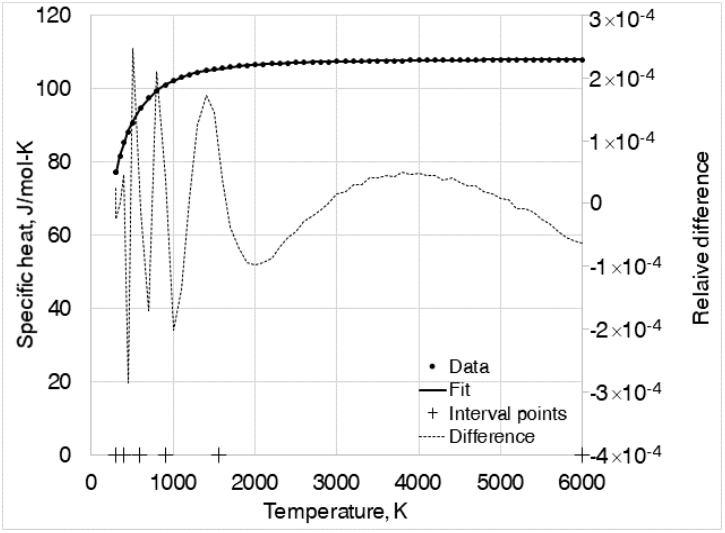


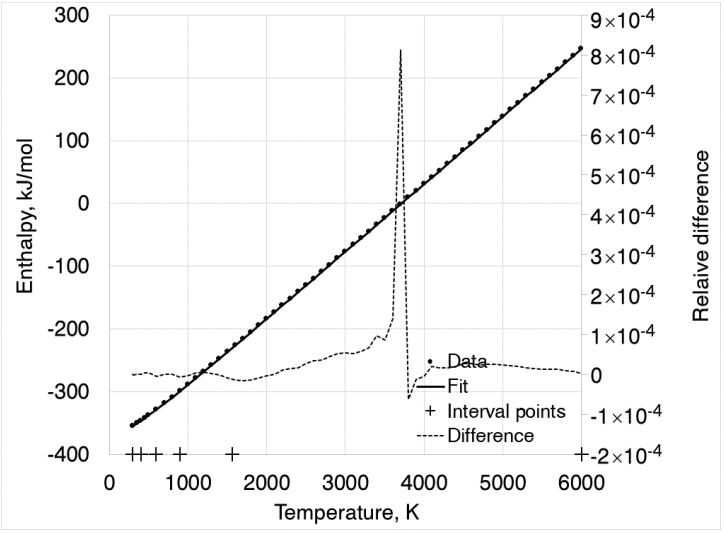


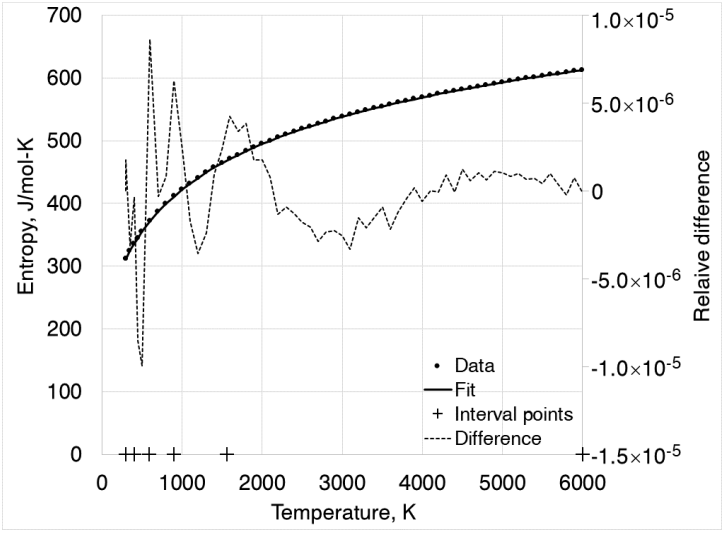


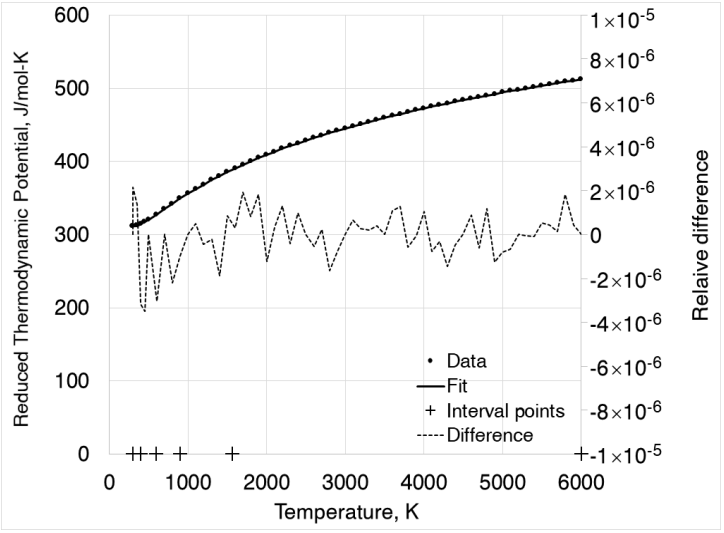
Sulfuryl Chloride (SO_2Cl_2) – Cl2O2S1(g)

MW = 134.96 g/mol – $\Delta_f H^o$(298.15 K) = -354.803 kJ/mol – S^o(298.15 K) = 311.095 J/mol-K – p^o = 0.1 MPa



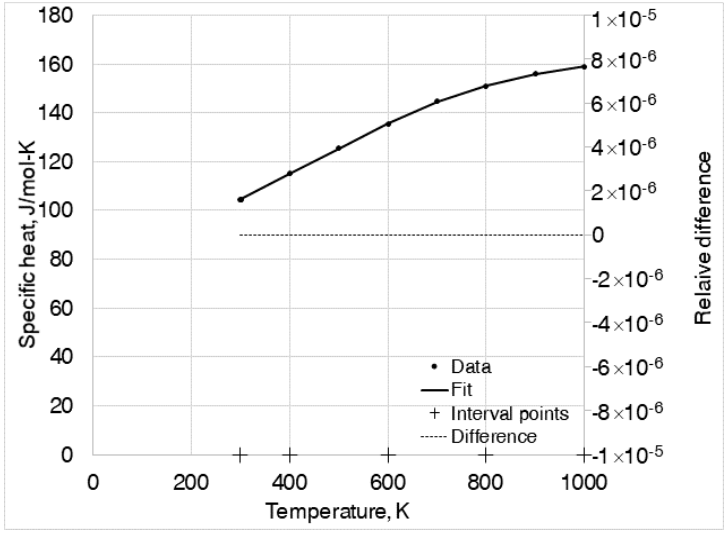


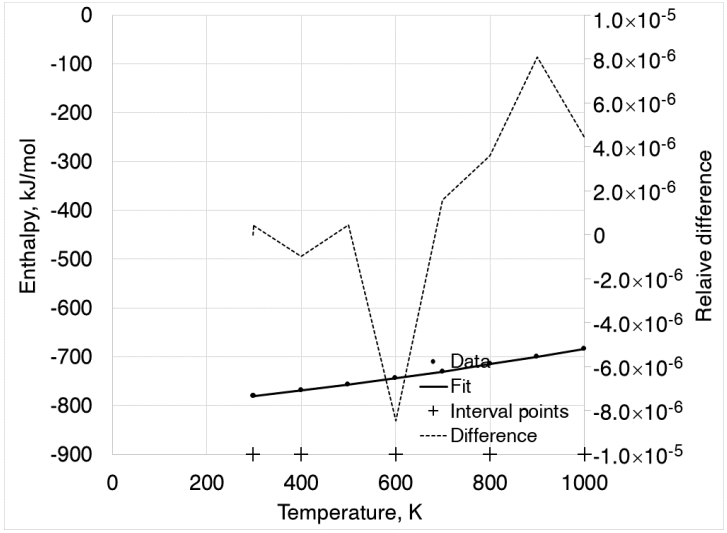


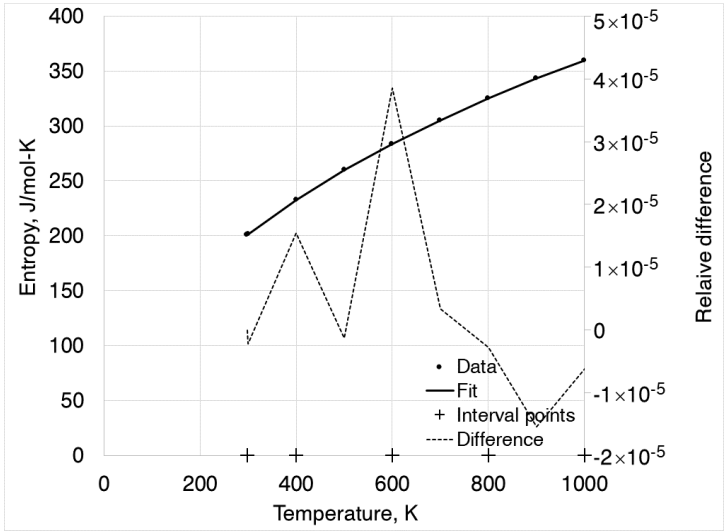


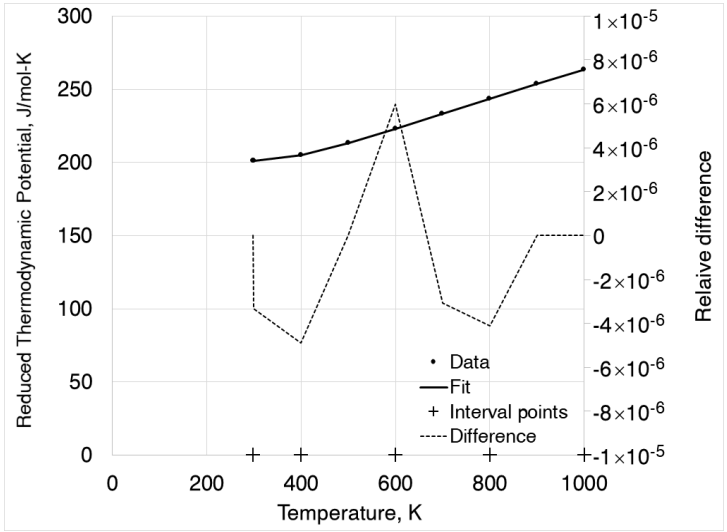
Tungsten Chloride Oxide (WO_2Cl_2) – Cl2O2W1(cr)

MW = 286.75 g/mol – $\Delta_f H^o$(298.15 K) = -780.316 kJ/mol – S^o(298.15 K) = 200.832 J/mol-K – p^o = 0.1 MPa



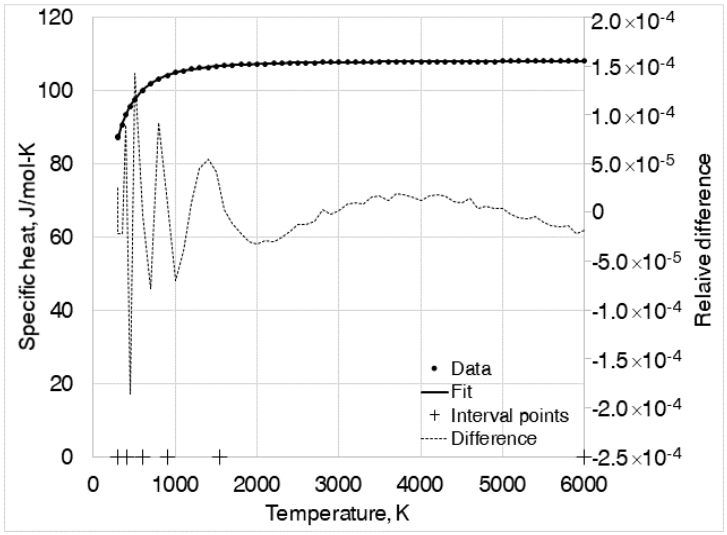


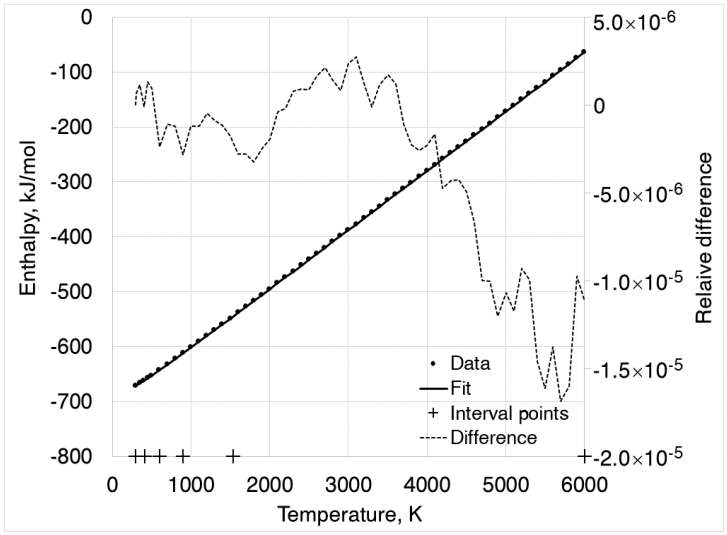


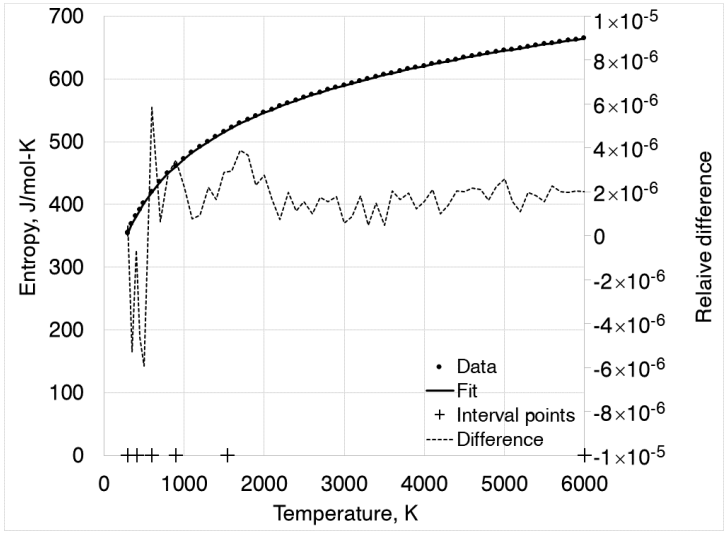


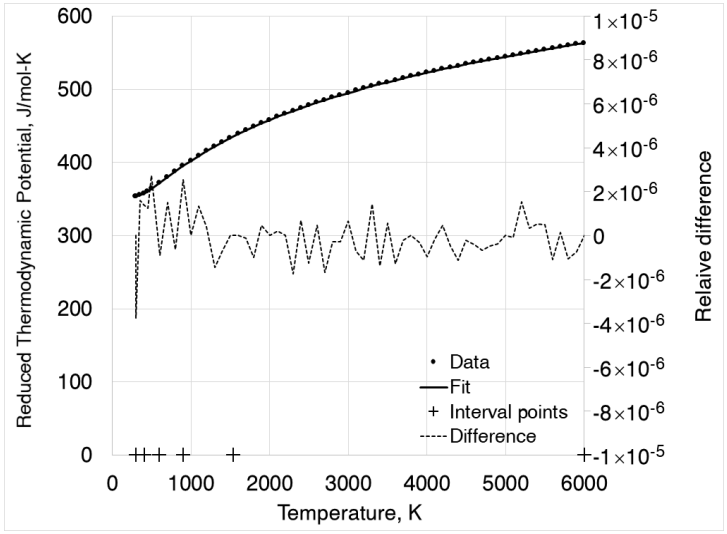
Tungsten Chloride Oxide (WO_2Cl_2) – Cl2O2W1(g)

MW = 286.75 g/mol – $\Delta_f H^o$(298.15 K) = -671.532 kJ/mol – S^o(298.15 K) = 353.927 J/mol-K – p^o = 0.1 MPa



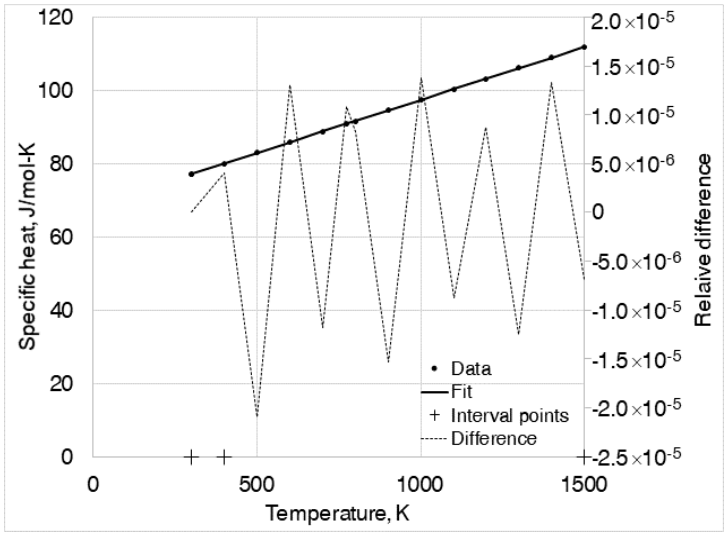


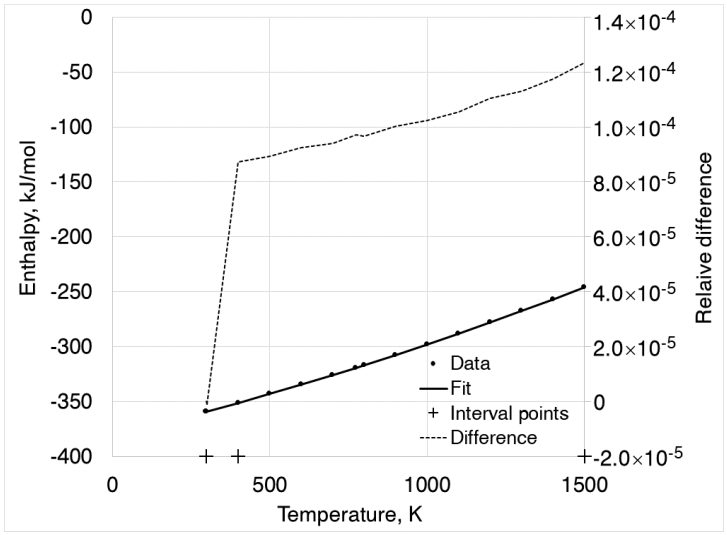


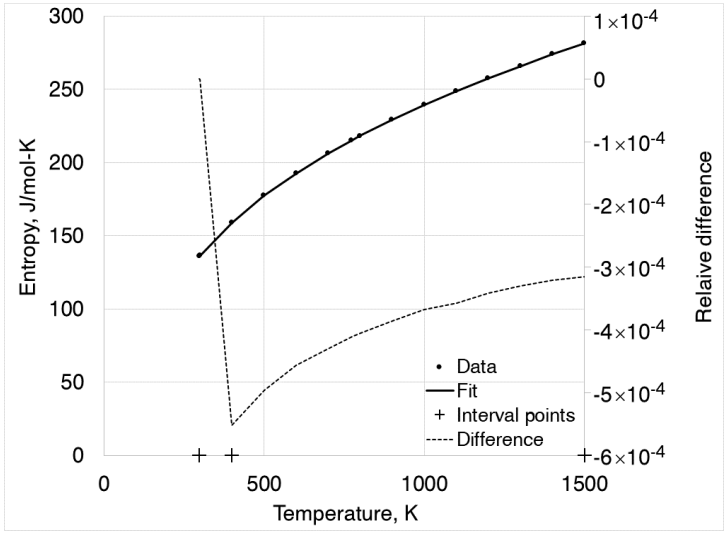


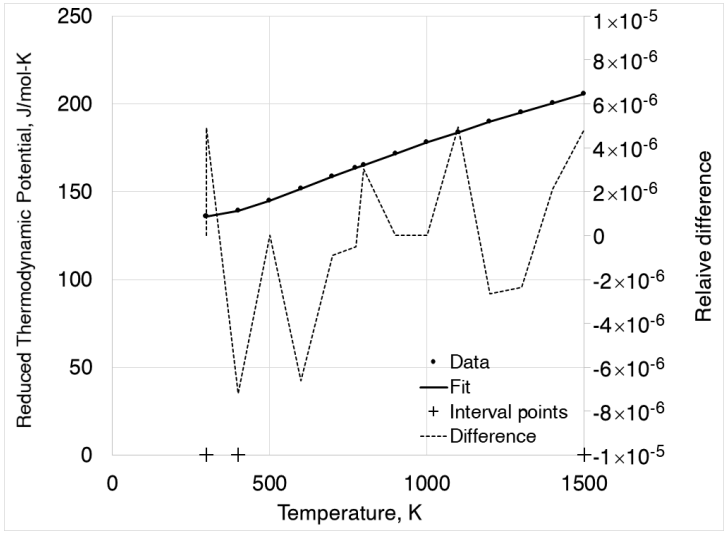
Lead Chloride ($PbCl_2$) – Cl2Pb1(cr)

MW = 278.1 g/mol – $\Delta_f H^o$(298.15 K) = -359.406 kJ/mol – S^o(298.15 K) = 135.98 J/mol-K – p^o = 0.1 MPa



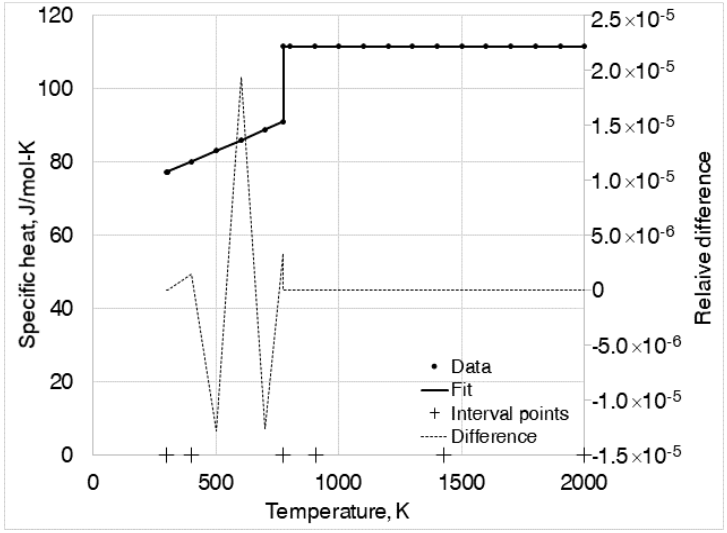


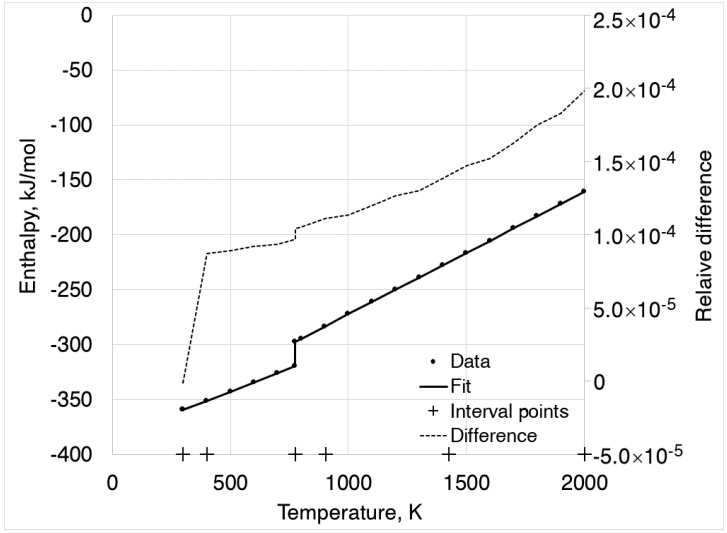


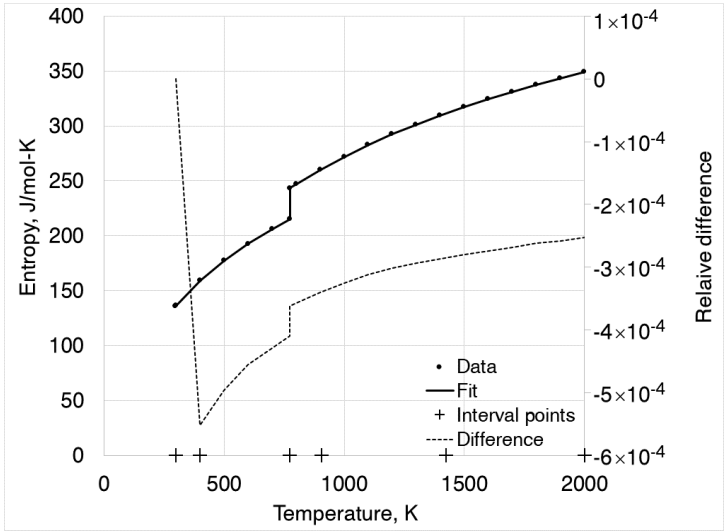


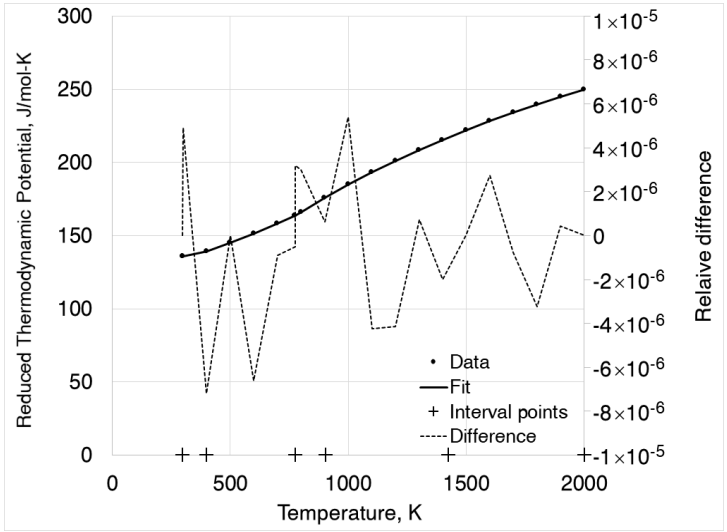
Lead Chloride ($PbCl_2$) – Cl2Pb1(cr,l)

MW = 278.1 g/mol – $\Delta_f H^o$(298.15 K) = -359.406 kJ/mol – S^o(298.15 K) = 135.98 J/mol-K – p^o = 0.1 MPa



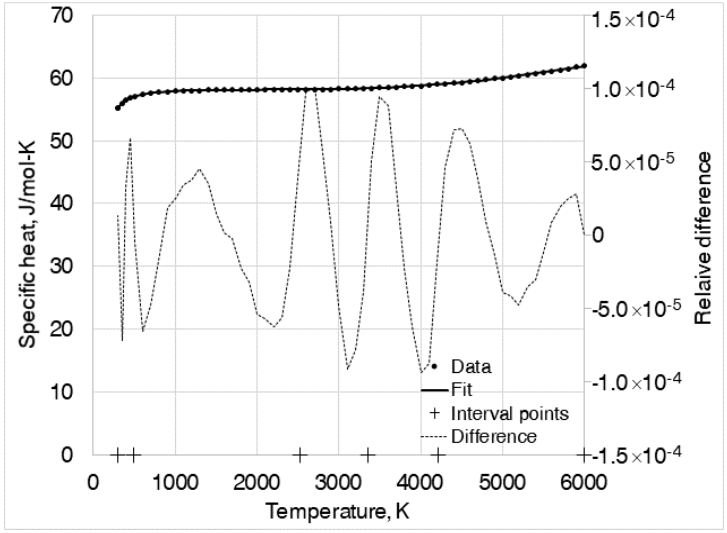


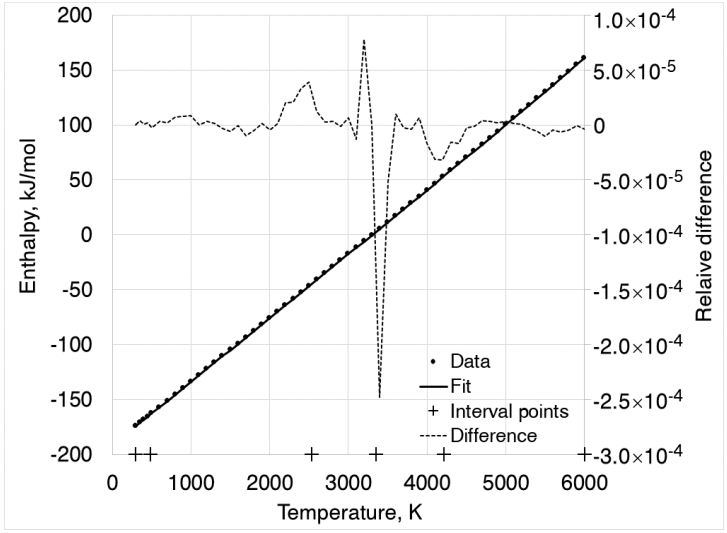


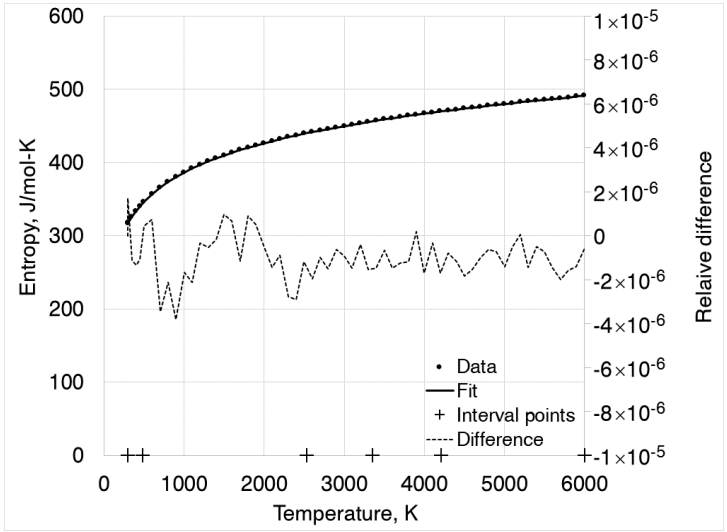


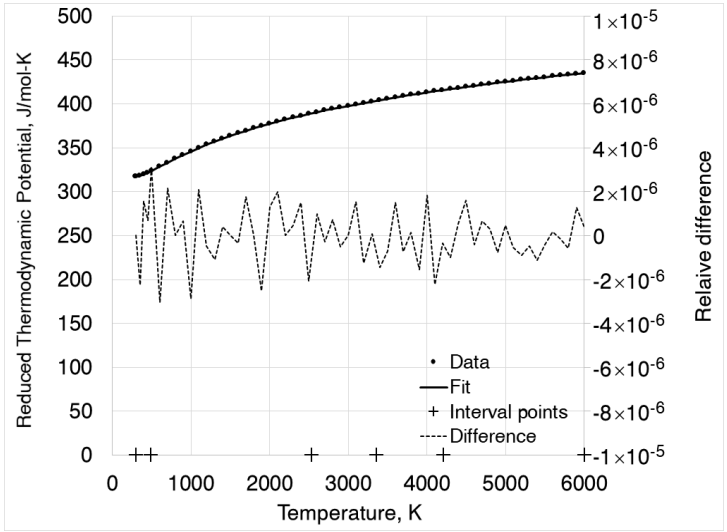
Lead Chloride ($PbCl_2$) – Cl2Pb1(g)

MW = 278.1 g/mol – $\Delta_f H^o$(298.15 K) = -174.054 kJ/mol – S^o(298.15 K) = 317.205 J/mol-K – p^o = 0.1 MPa



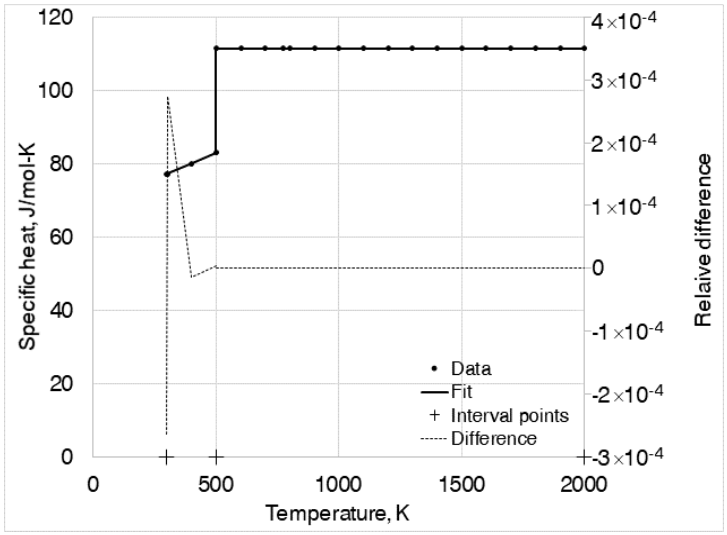


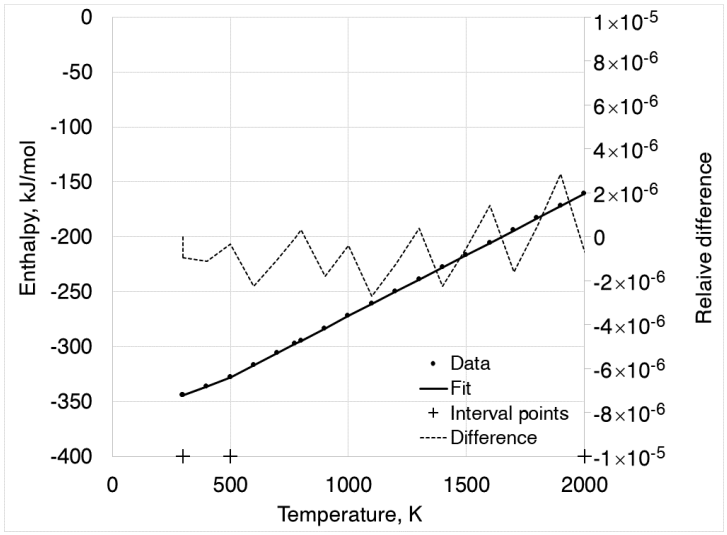


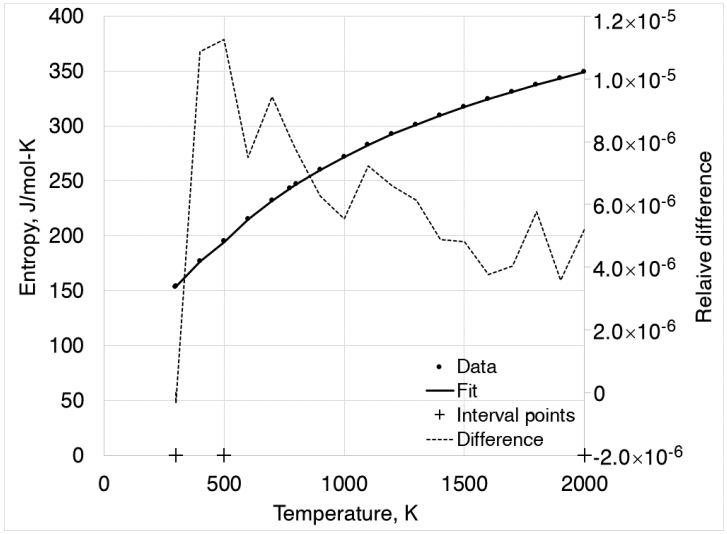


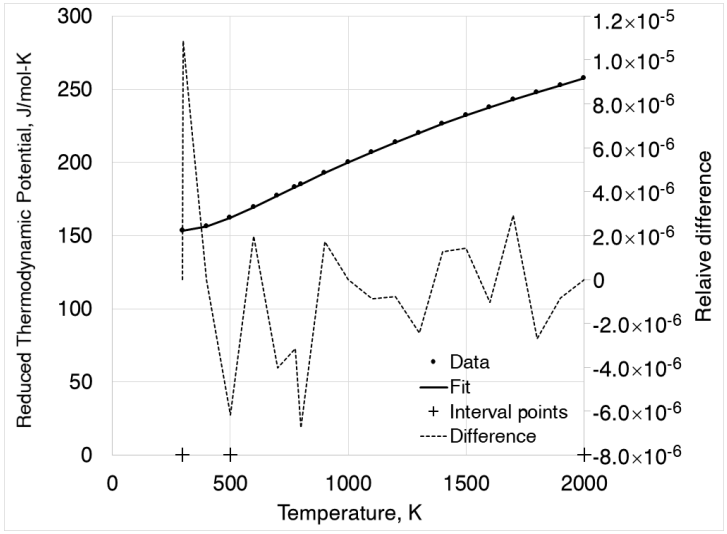
Lead Chloride ($PbCl_2$) – Cl2Pb1(l)

MW = 278.1 g/mol – $\Delta_f H^o$(298.15 K) = -344.251 kJ/mol – S^o(298.15 K) = 153.396 J/mol-K – p^o = 0.1 MPa



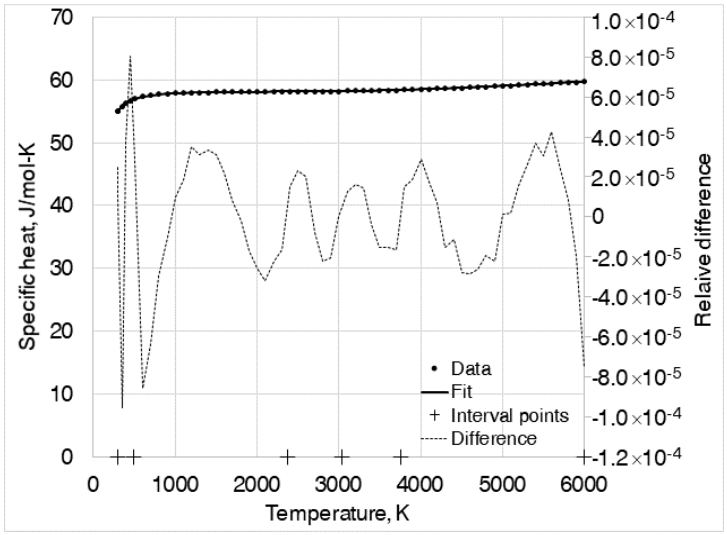


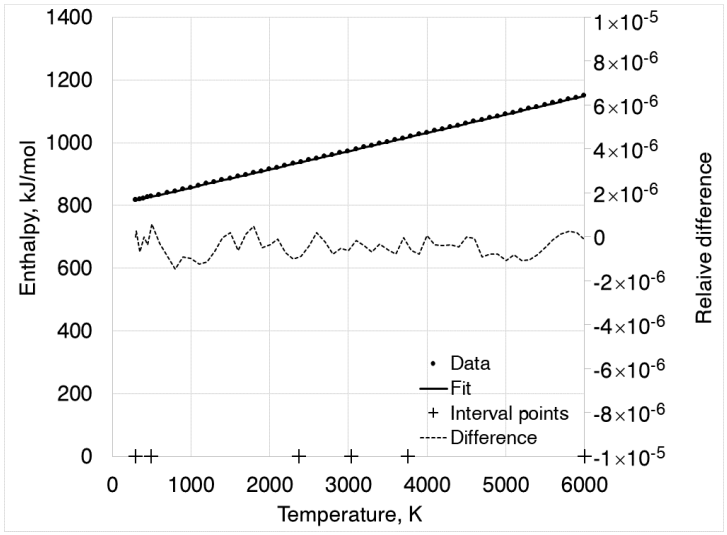


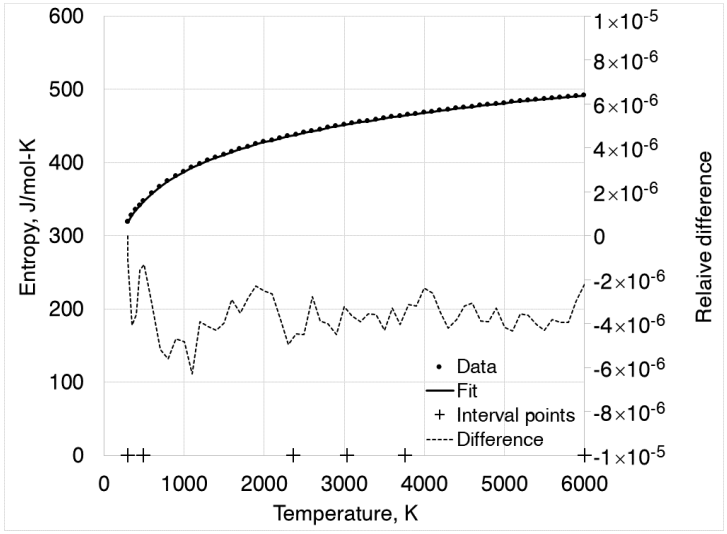


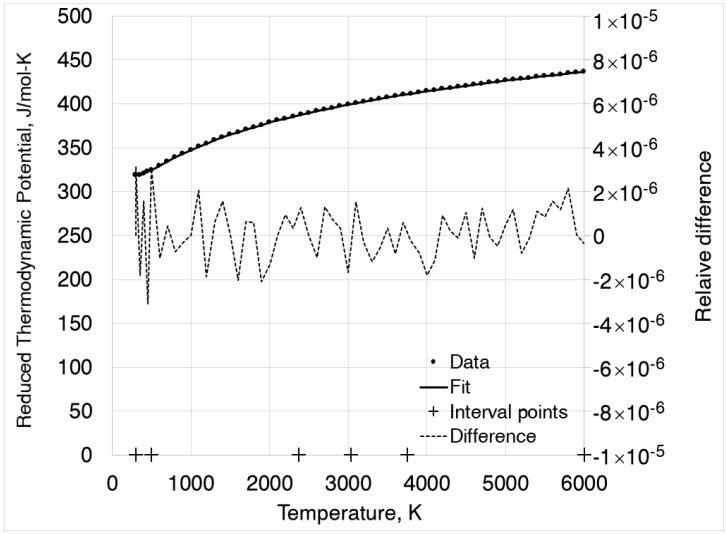
Lead Chloride, Ion ($PbCl_2^+$) – Cl2Pb1+(g)

MW = 278.1 g/mol – $\Delta_f H^o$(298.15 K) = 816.298 kJ/mol – S^o(298.15 K) = 318.566 J/mol-K – p^o = 0.1 MPa



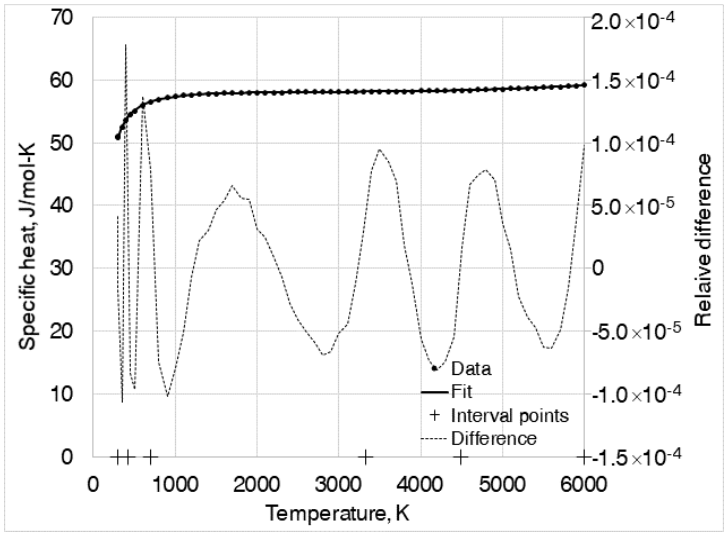


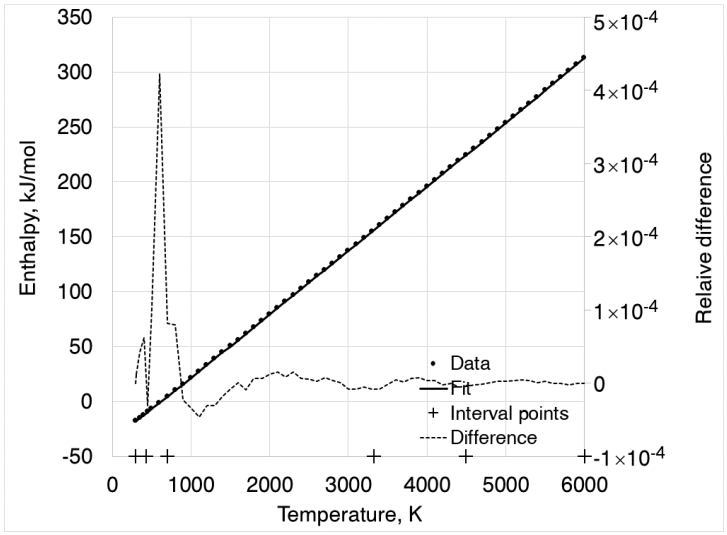


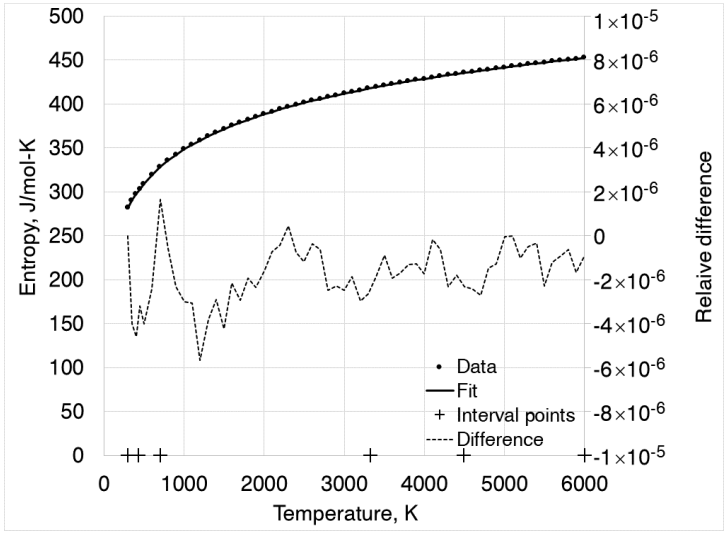


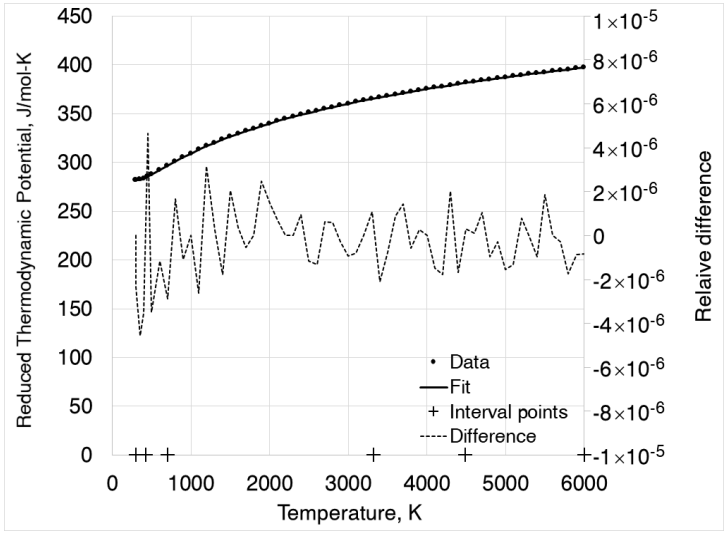
Sulfur Chloride (SCl_2) – Cl2S1(g)

MW = 102.97 g/mol – $\Delta_f H^o$(298.15 K) = -17.573 kJ/mol – S^o(298.15 K) = 281.628 J/mol-K – p^o = 0.1 MPa



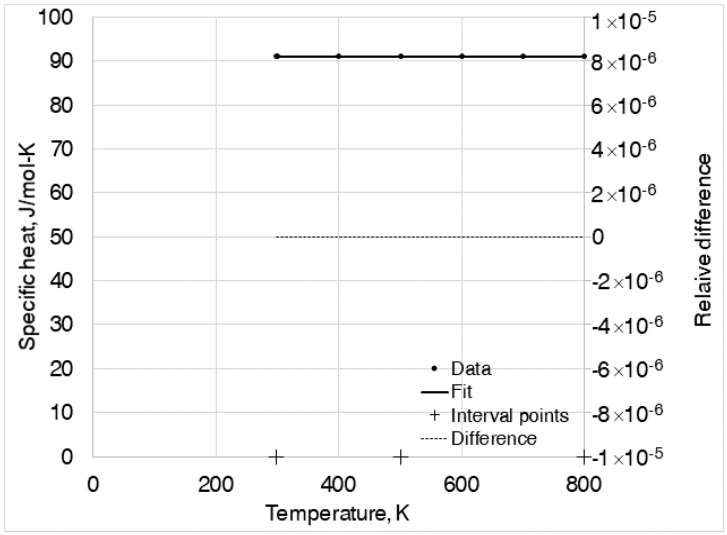


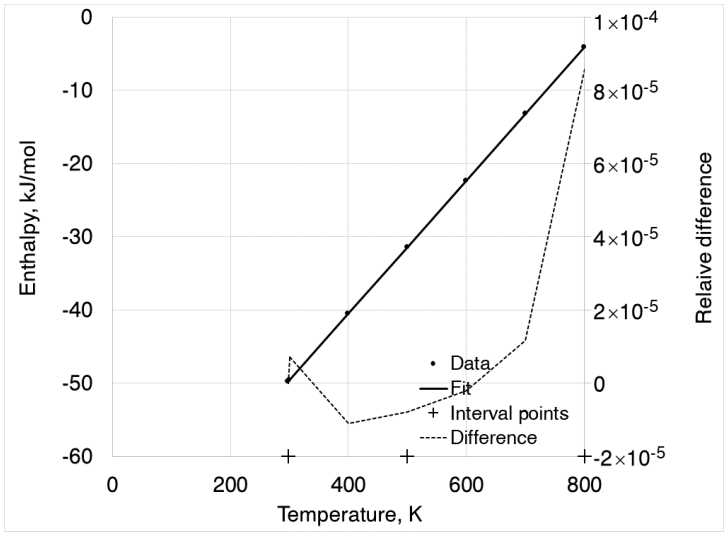


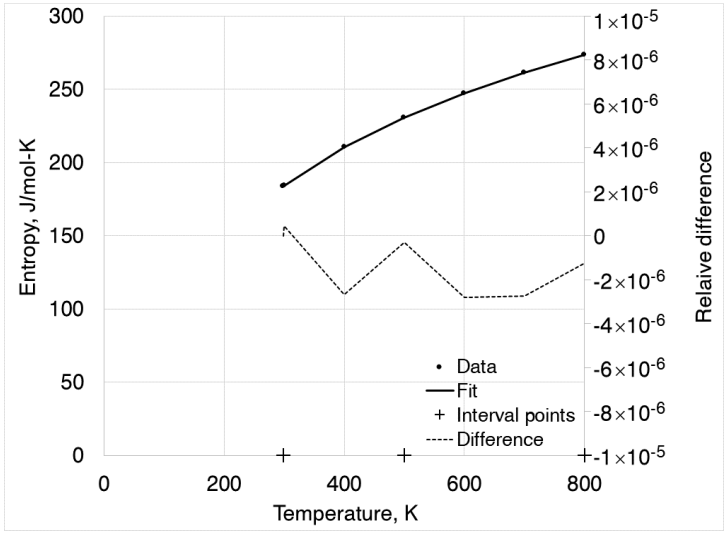


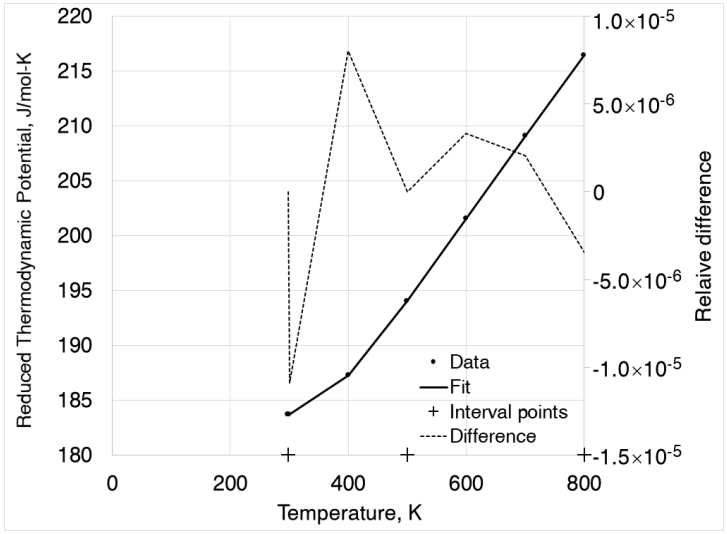
Sulfur Chloride (SCl_2) – Cl2S1(l)

MW = 102.97 g/mol – $\Delta_f H^o$(298.15 K) = -49.79 kJ/mol – S^o(298.15 K) = 183.678 J/mol-K – p^o = 0.1 MPa



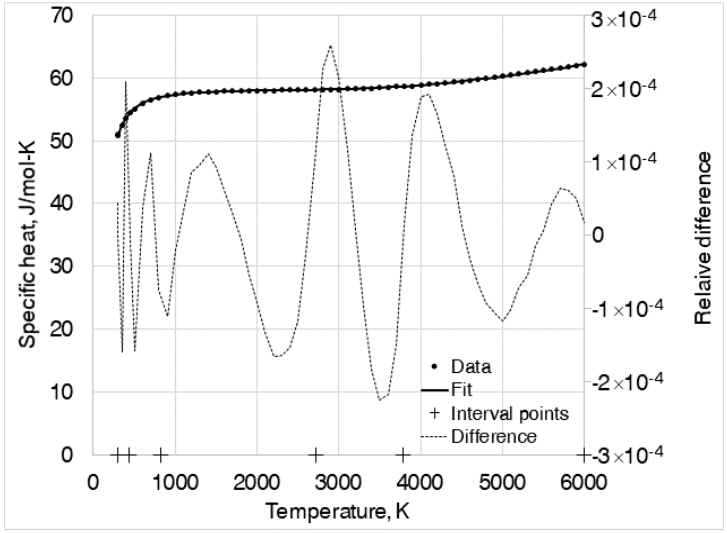


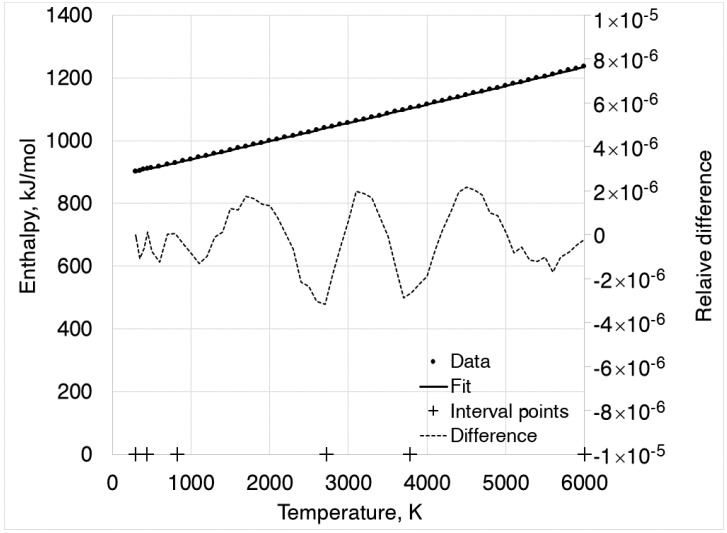


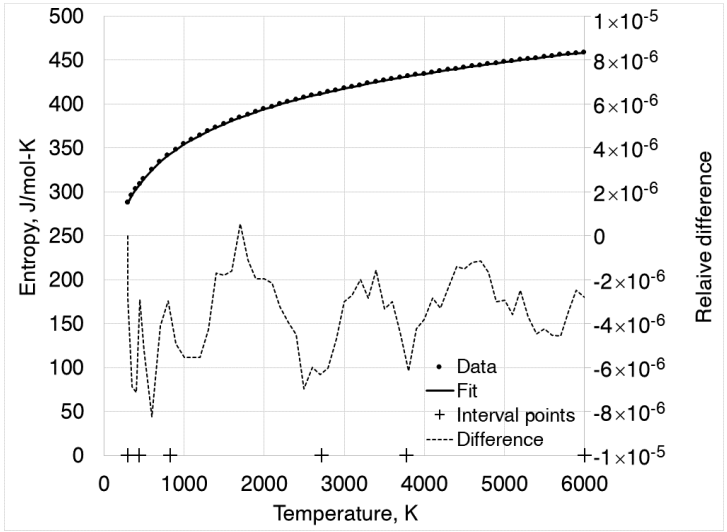


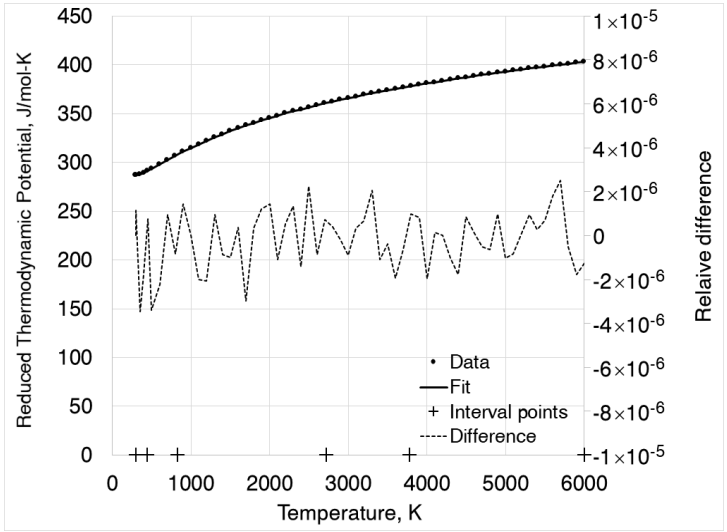
Sulfur Chloride, Ion (SCl_2^+) – Cl2S1+(g)

MW = 102.97 g/mol – $\Delta_f H^o$(298.15 K) = 901.485 kJ/mol – S^o(298.15 K) = 287.322 J/mol-K – p^o = 0.1 MPa



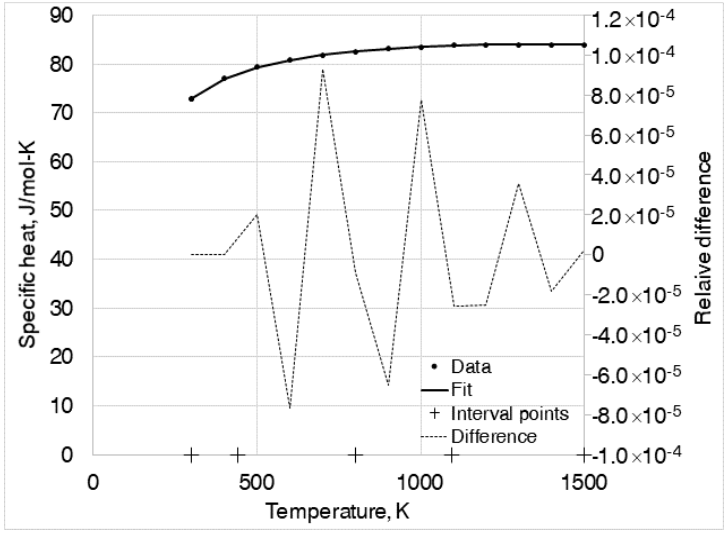


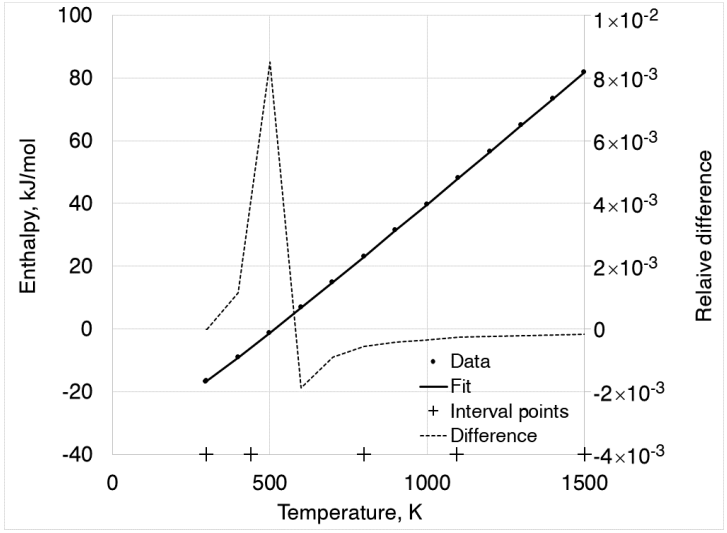


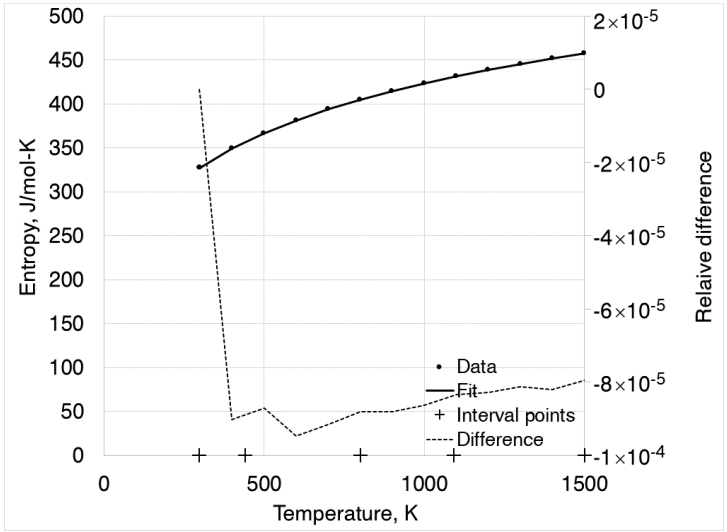


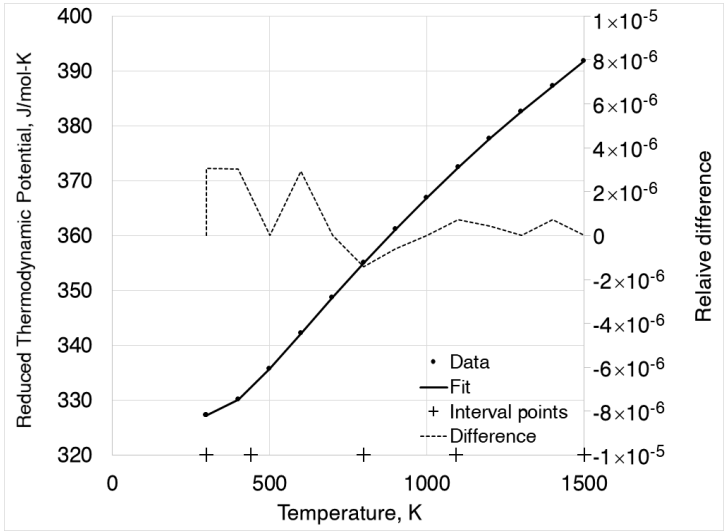
Sulfur Chloride (ClSSCl) – Cl2S2(g)

MW = 135.03 g/mol – $\Delta_f H^o$(298.15 K) = -16.736 kJ/mol – S^o(298.15 K) = 327.215 J/mol-K – p^o = 0.1 MPa



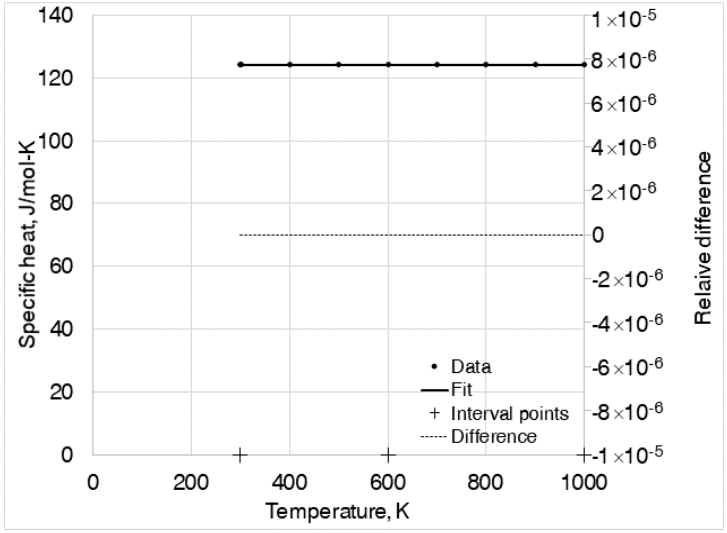


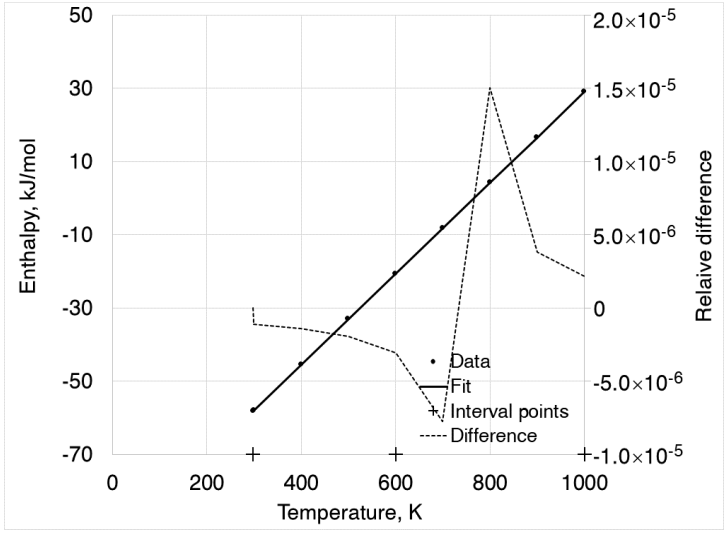


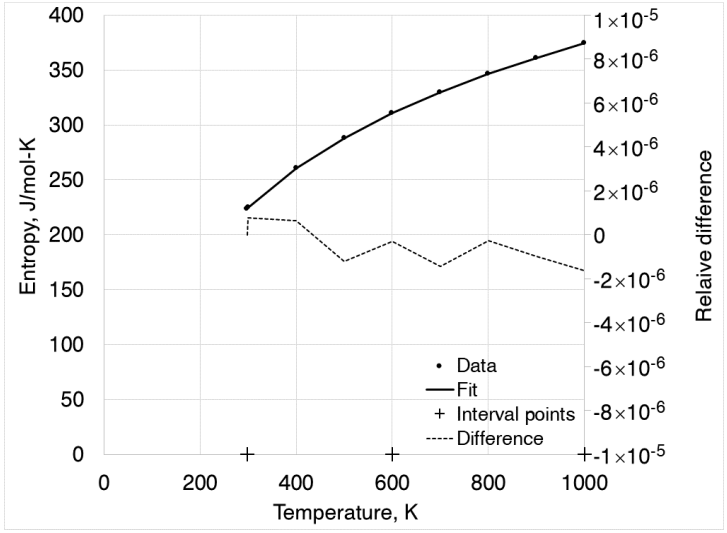


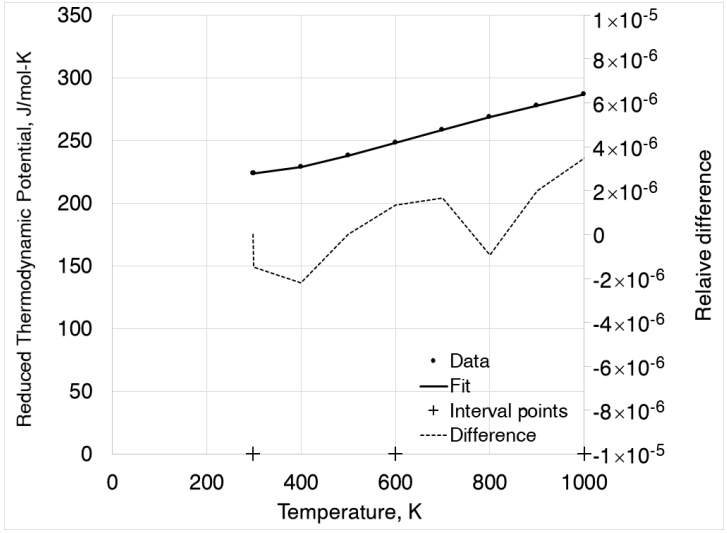
Sulfur Chloride (ClSSCl) – Cl2S2(l)

MW = 135.03 g/mol – $\Delta_f H^o$(298.15 K) = -58.158 kJ/mol – S^o(298.15 K) = 223.844 J/mol-K – p^o = 0.1 MPa



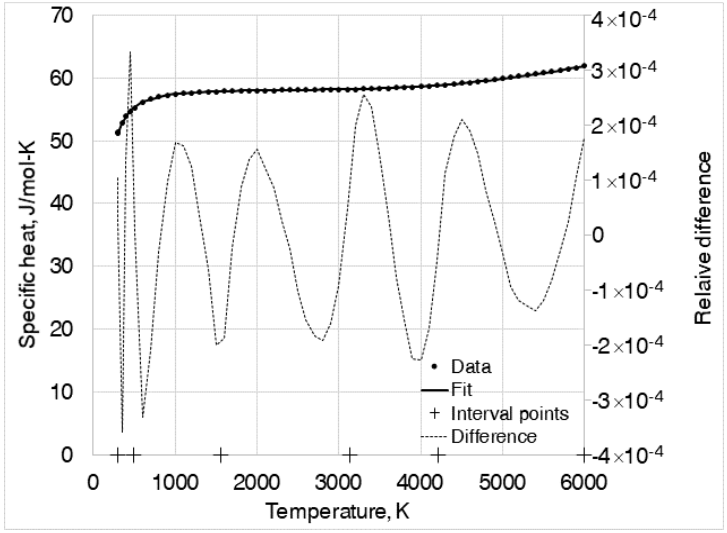


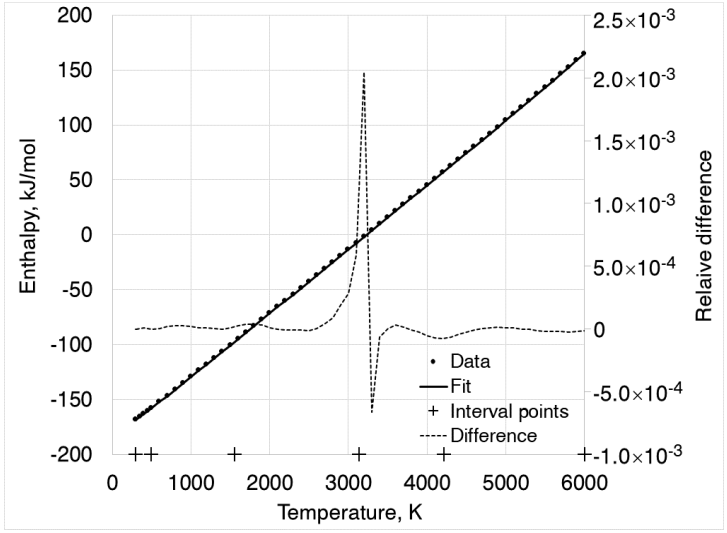


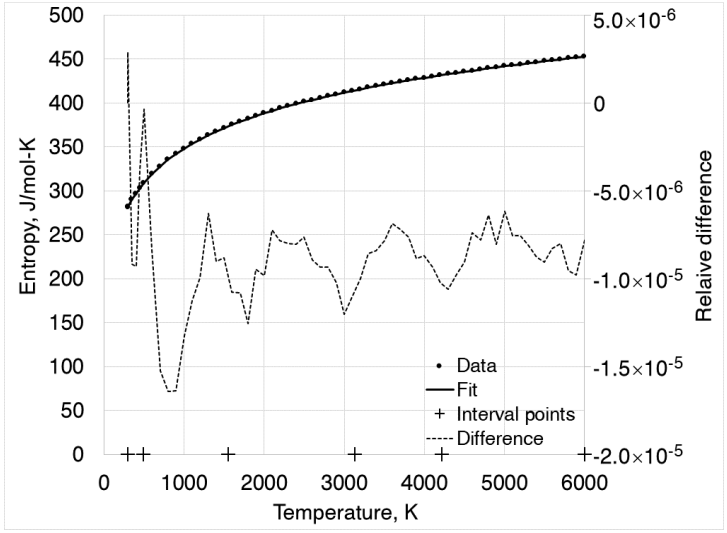


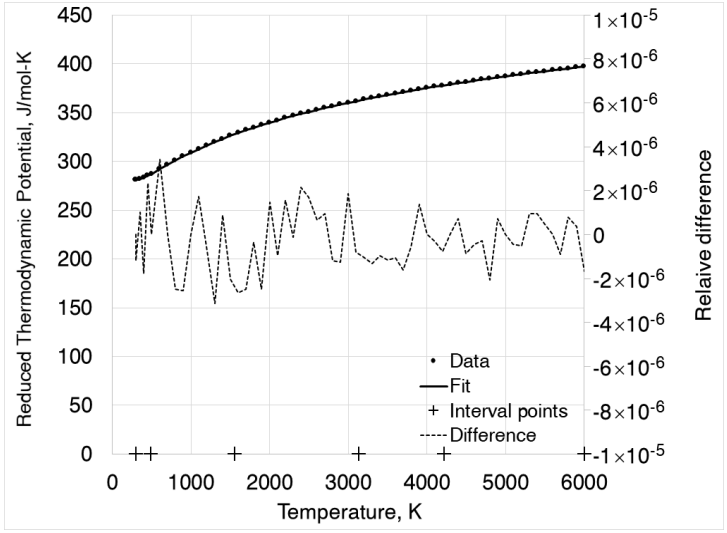
Dichlorosilylene ($SiCl_2$) – Cl2Si1(g)

MW = 98.992 g/mol – $\Delta_f H^o$(298.15 K) = -168.615 kJ/mol – S^o(298.15 K) = 281.333 J/mol-K – p^o = 0.1 MPa



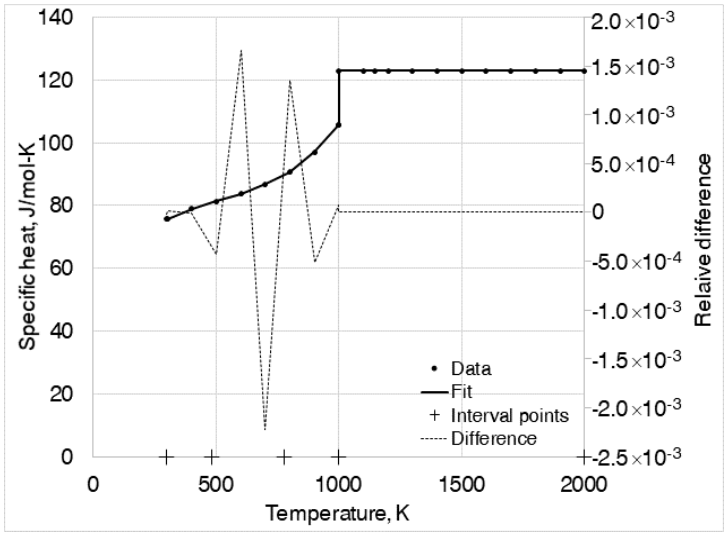


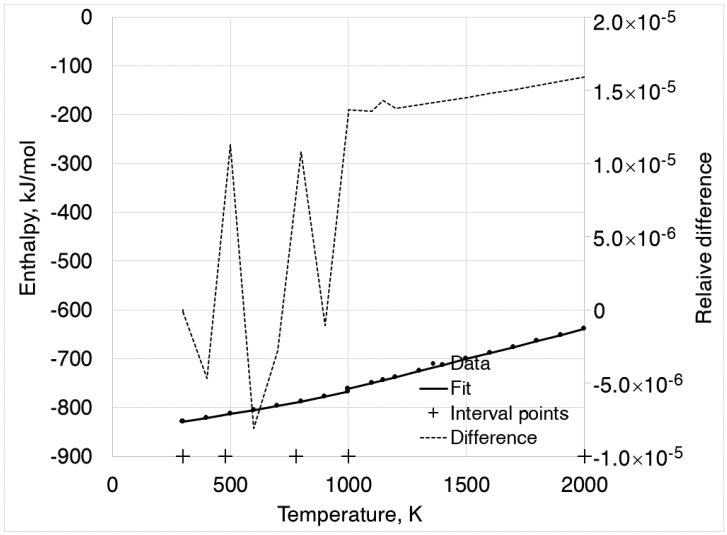


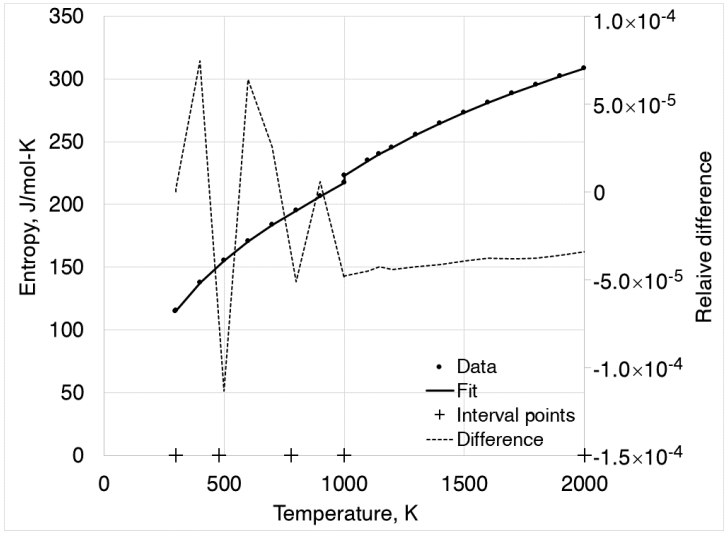


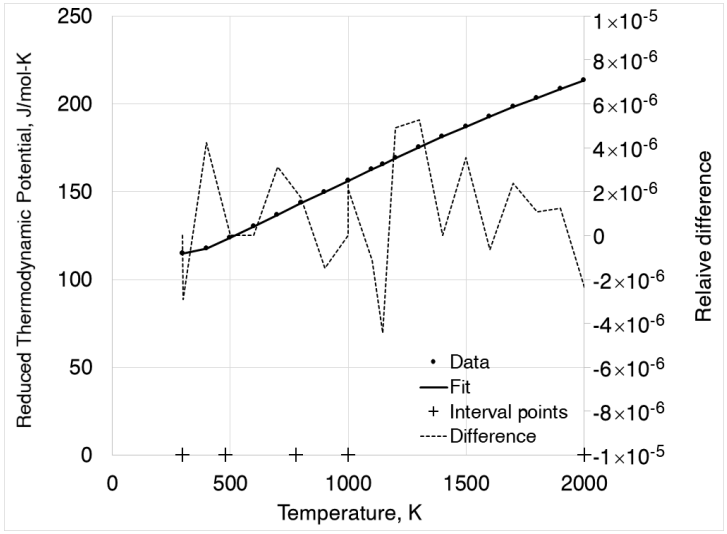
Strontium Chloride ($SrCl_2$) – Cl2Sr1(cr)

MW = 158.53 g/mol – $\Delta_f H^o$(298.15 K) = -828.85 kJ/mol – S^o(298.15 K) = 114.809 J/mol-K – p^o = 0.1 MPa



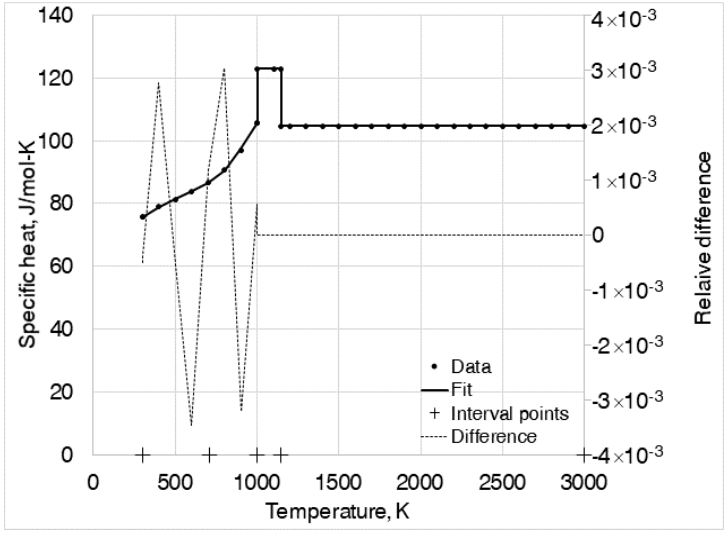


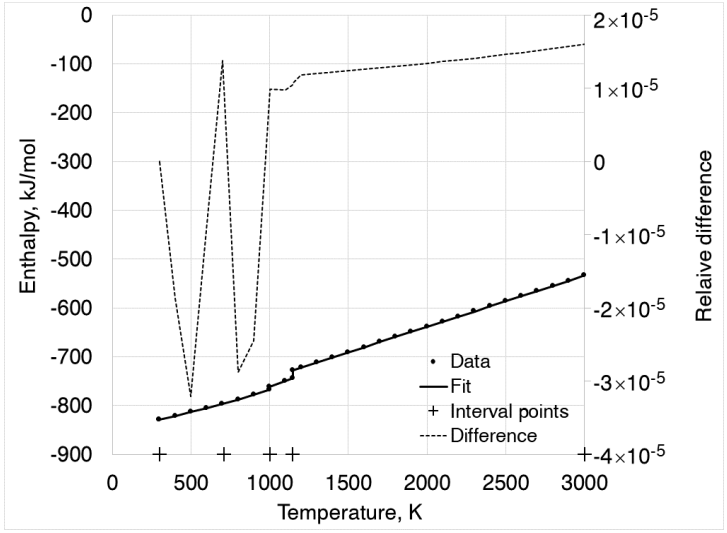


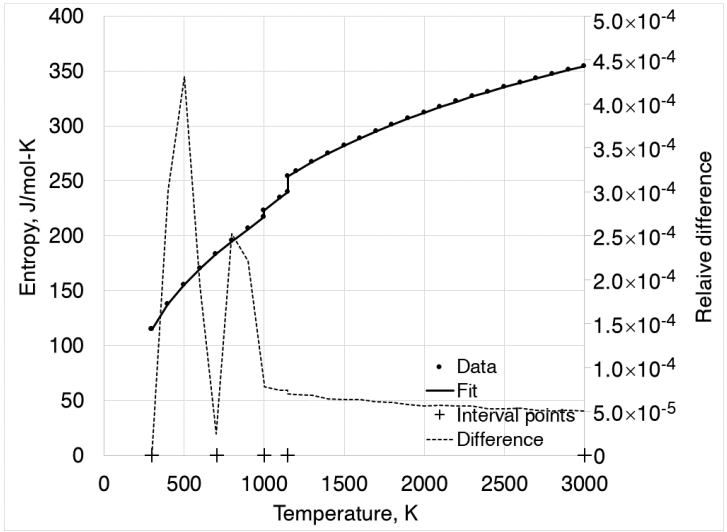


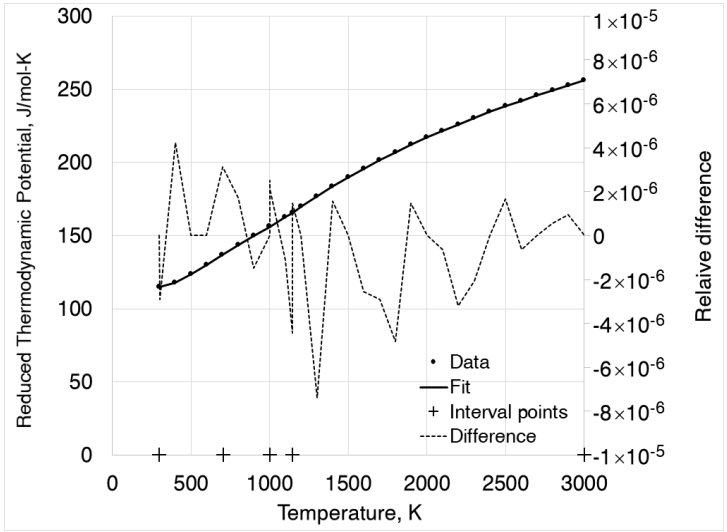
Strontium Chloride ($SrCl_2$) – Cl2Sr1(cr,l)

MW = 158.53 g/mol – $\Delta_f H^o$(298.15 K) = -828.85 kJ/mol – S^o(298.15 K) = 114.809 J/mol-K – p^o = 0.1 MPa



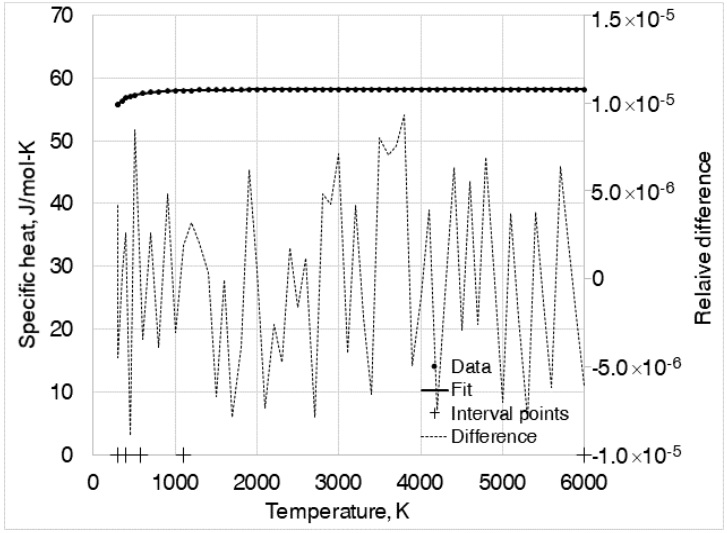


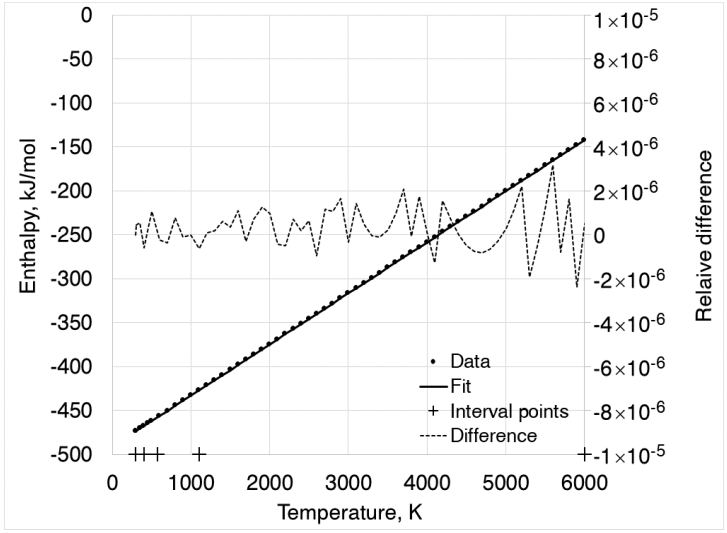


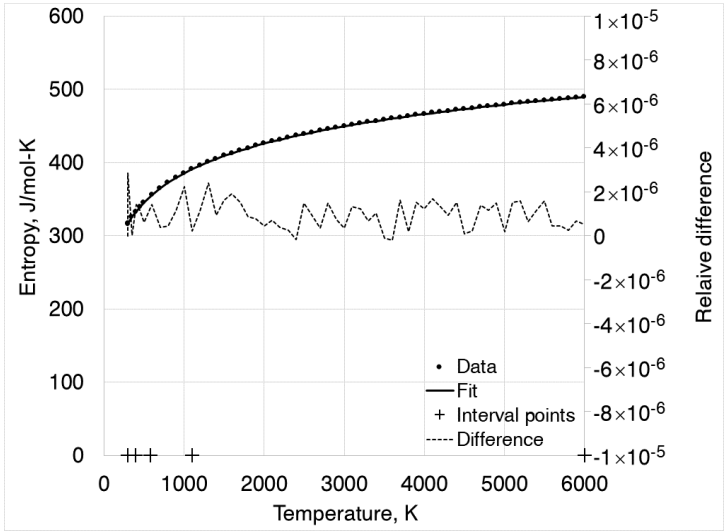


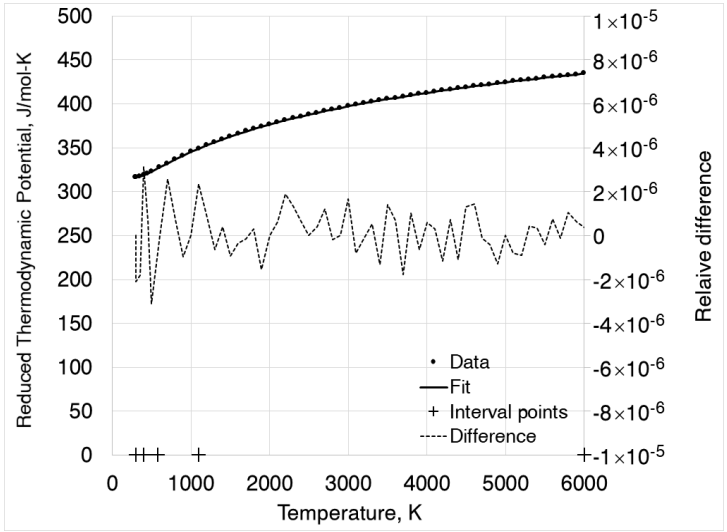
Strontium Chloride ($SrCl_2$) – Cl2Sr1(g)

MW = 158.53 g/mol – $\Delta_f H^o$(298.15 K) = -473.21 kJ/mol – S^o(298.15 K) = 316.339 J/mol-K – p^o = 0.1 MPa



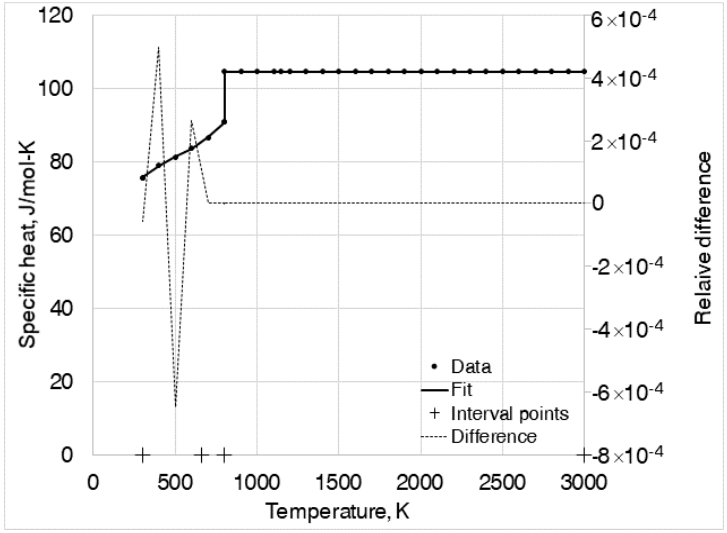


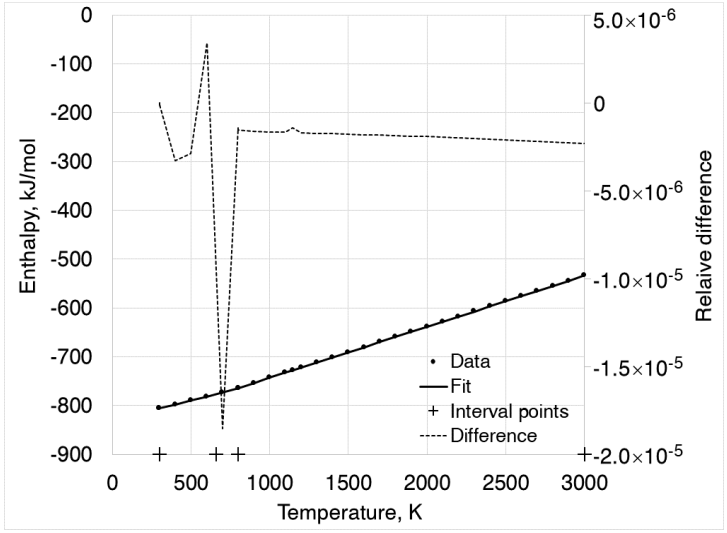


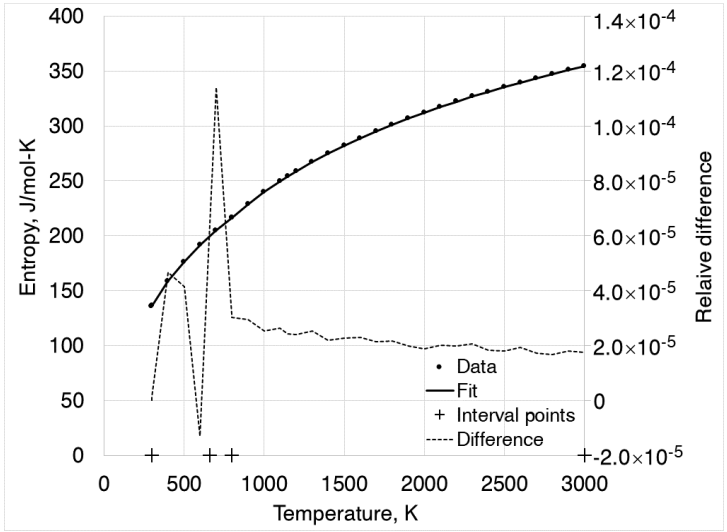


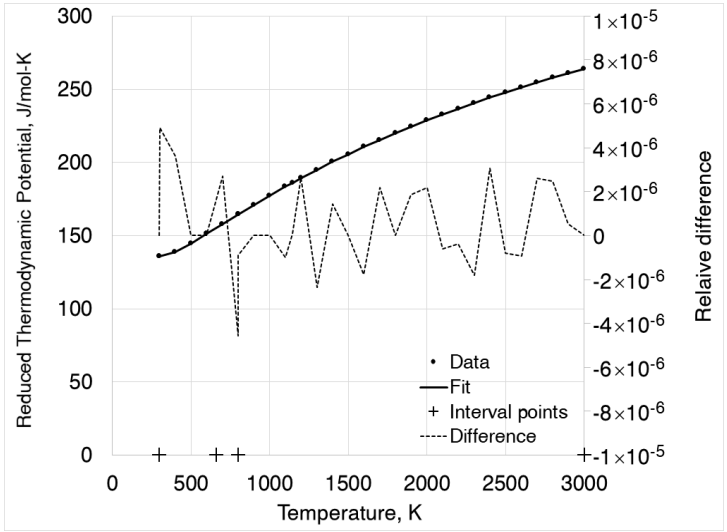
Strontium Chloride ($SrCl_2$) – Cl2Sr1(l)

MW = 158.53 g/mol – $\Delta_f H^o$(298.15 K) = -805.351 kJ/mol – S^o(298.15 K) = 135.82 J/mol-K – p^o = 0.1 MPa



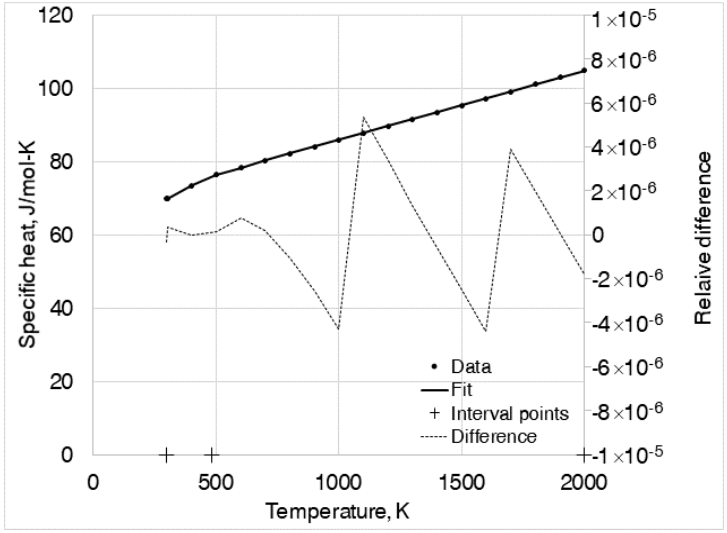


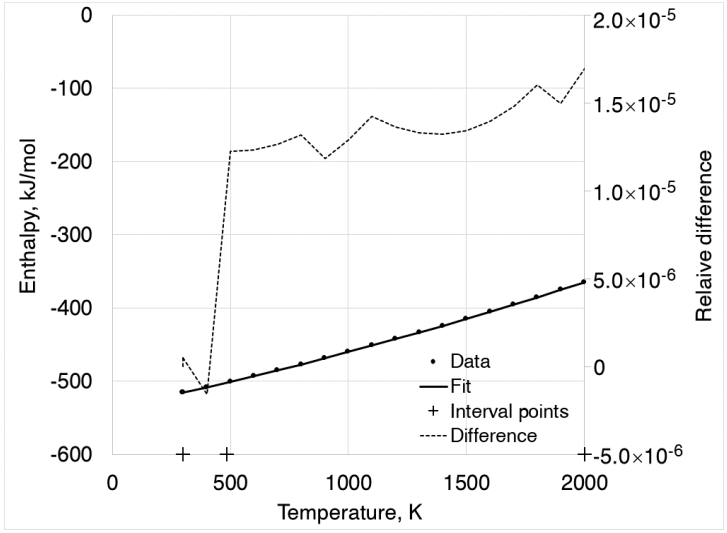


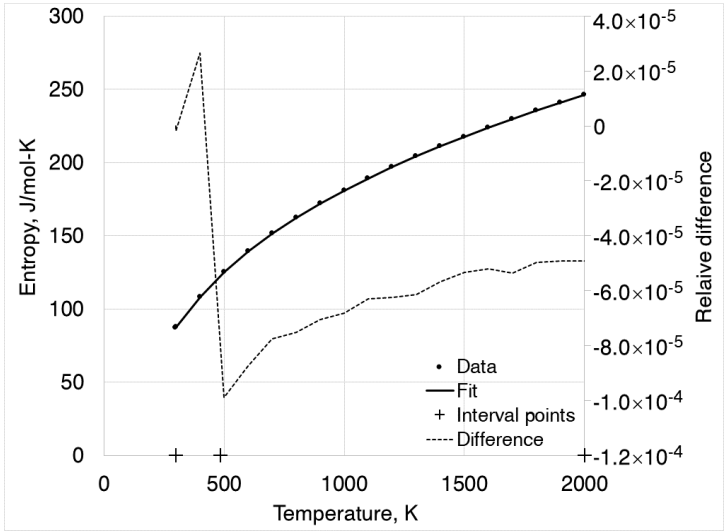


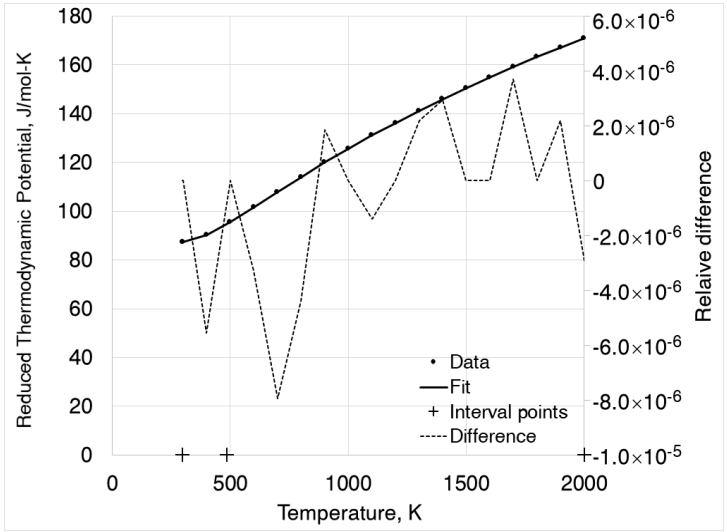
Titanium Chloride ($TiCl_2$) – Cl2Ti1(cr)

MW = 118.79 g/mol – $\Delta_f H^o$(298.15 K) = -515.469 kJ/mol – S^o(298.15 K) = 87.345 J/mol-K – p^o = 0.1 MPa



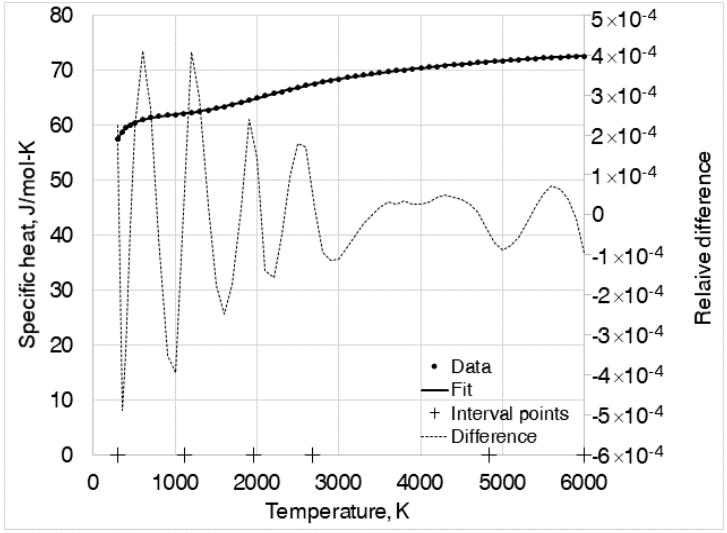


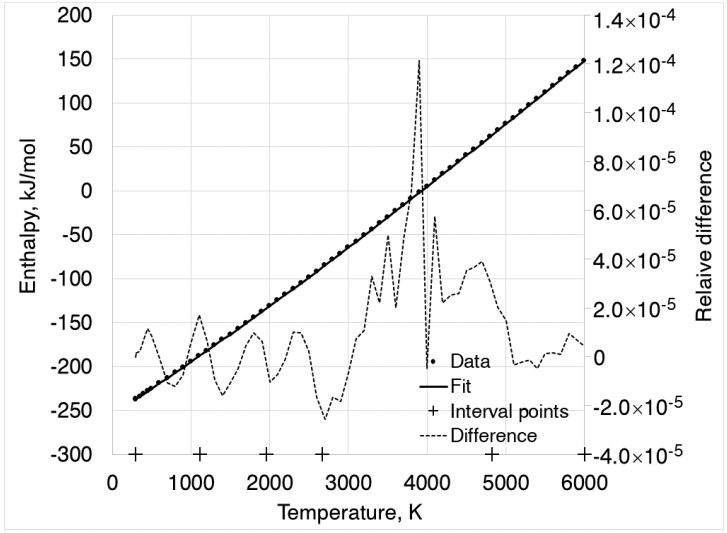


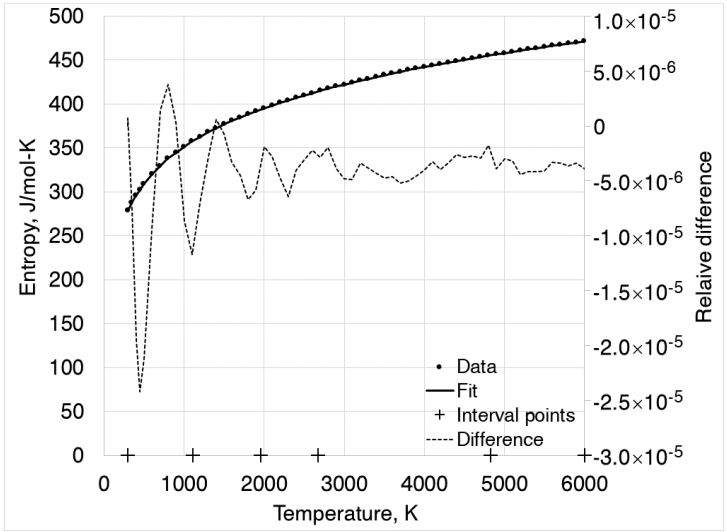


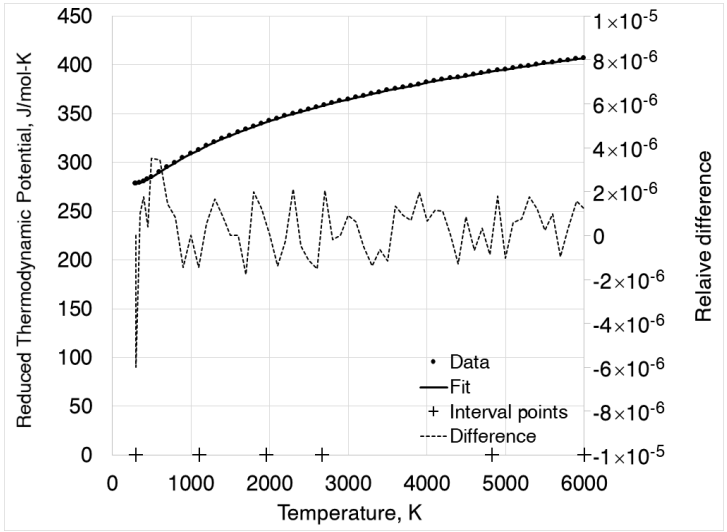
Titanium Chloride ($TiCl_2$) – Cl2Ti1(g)

MW = 118.79 g/mol – $\Delta_f H^o$(298.15 K) = -237.233 kJ/mol – S^o(298.15 K) = 278.338 J/mol-K – p^o = 0.1 MPa



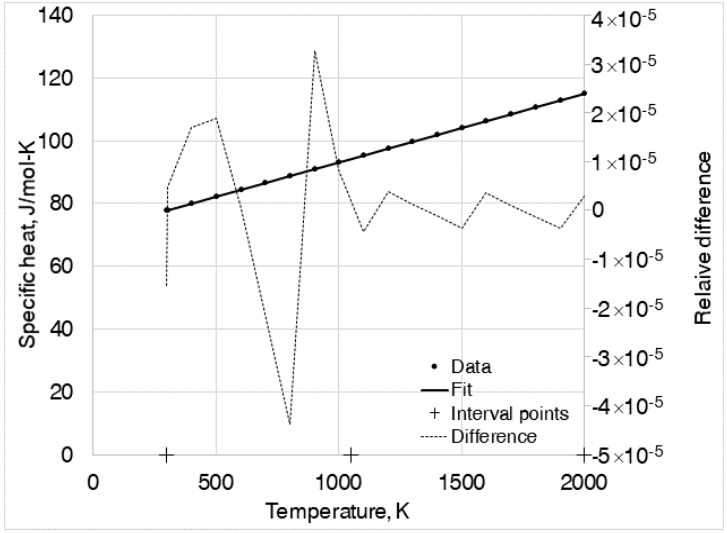


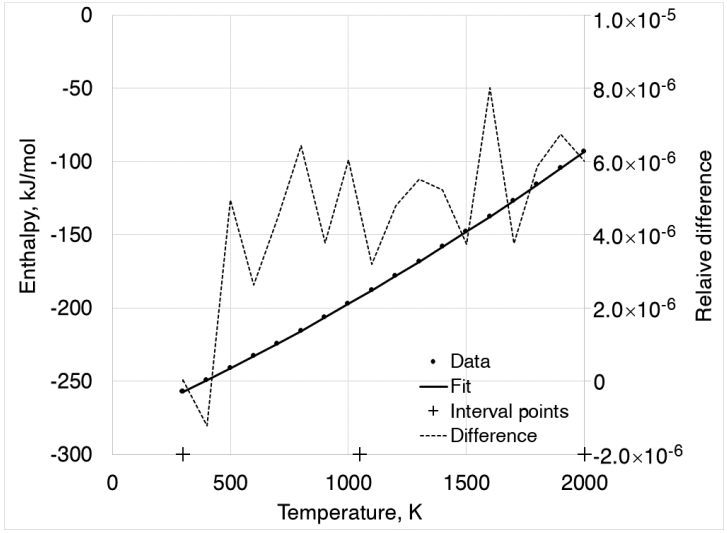


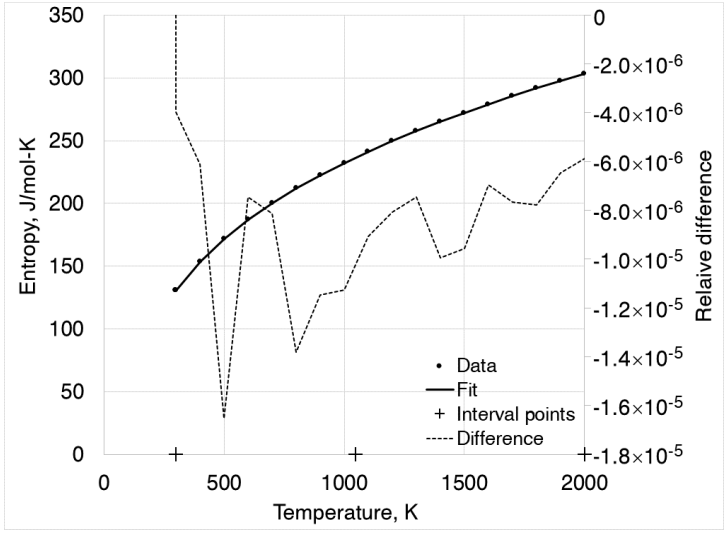


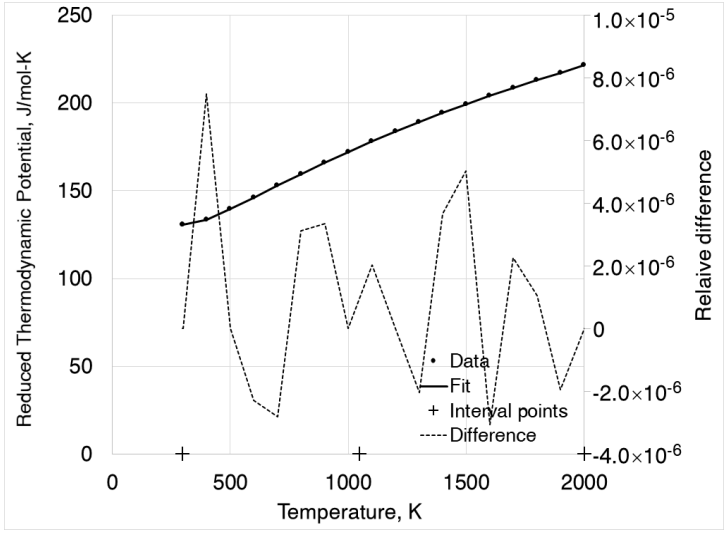
Tungsten Chloride (WCl_2) – Cl2W1(cr)

MW = 254.76 g/mol – $\Delta_f H^o$(298.15 K) = -257.316 kJ/mol – S^o(298.15 K) = 130.541 J/mol-K – p^o = 0.1 MPa



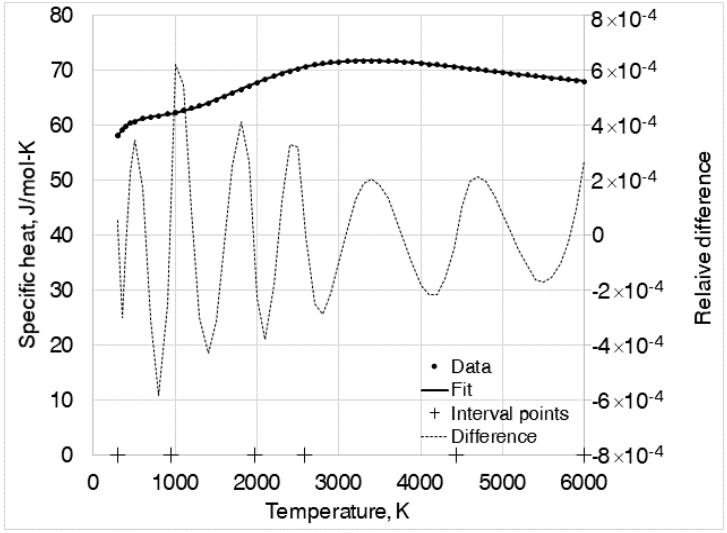


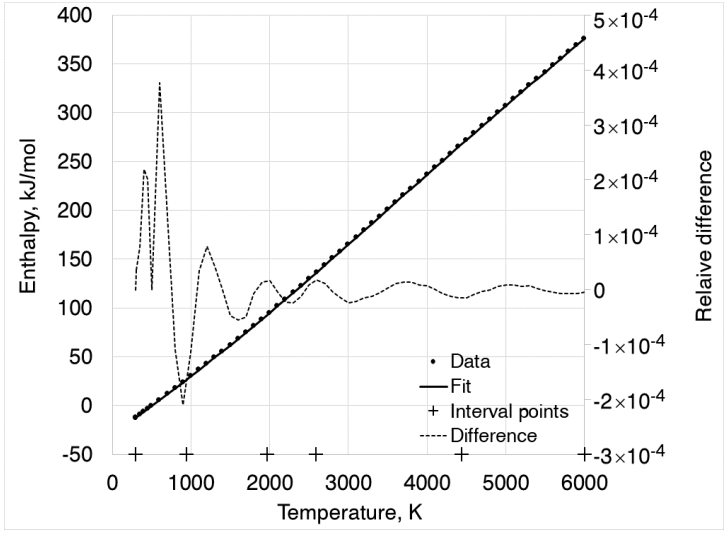


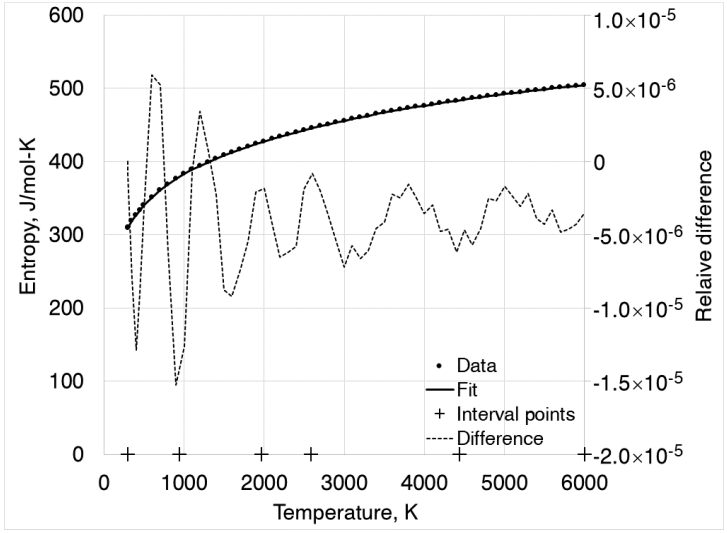


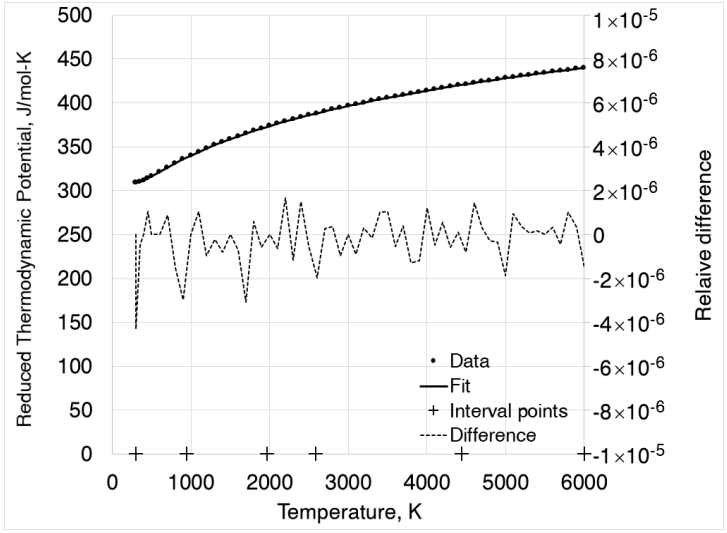
Tungsten Chloride (WCl_2) – Cl2W1(g)

MW = 254.76 g/mol – $\Delta_f H^o$(298.15 K) = -12.552 kJ/mol – S^o(298.15 K) = 309.416 J/mol-K – p^o = 0.1 MPa



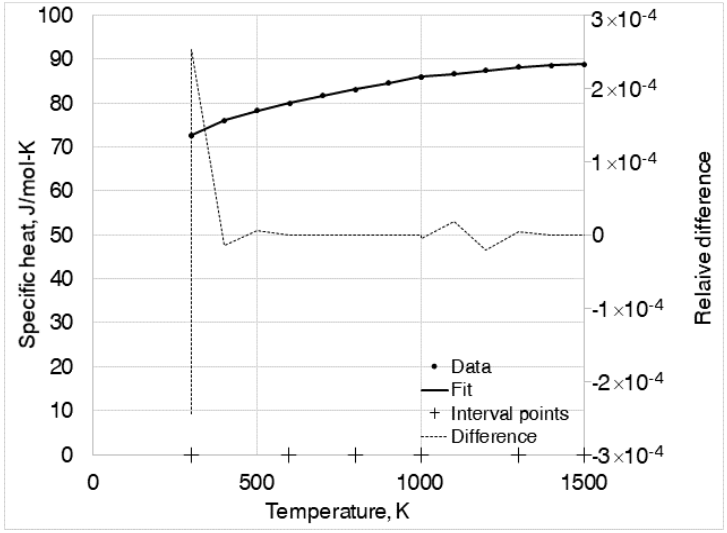


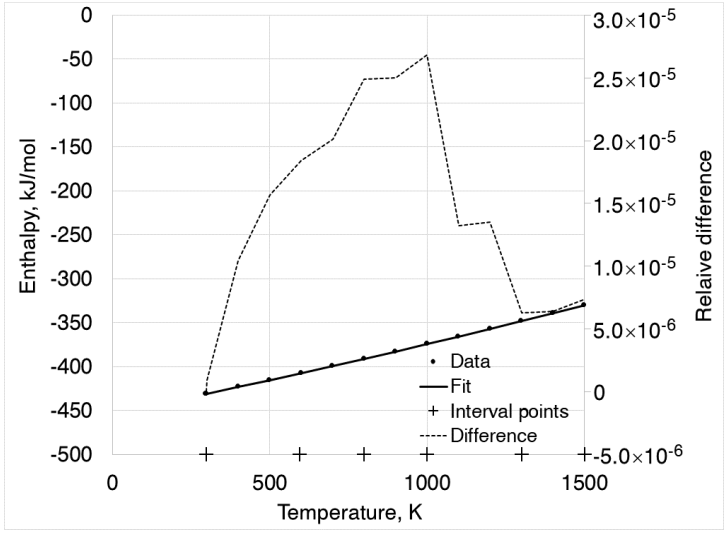


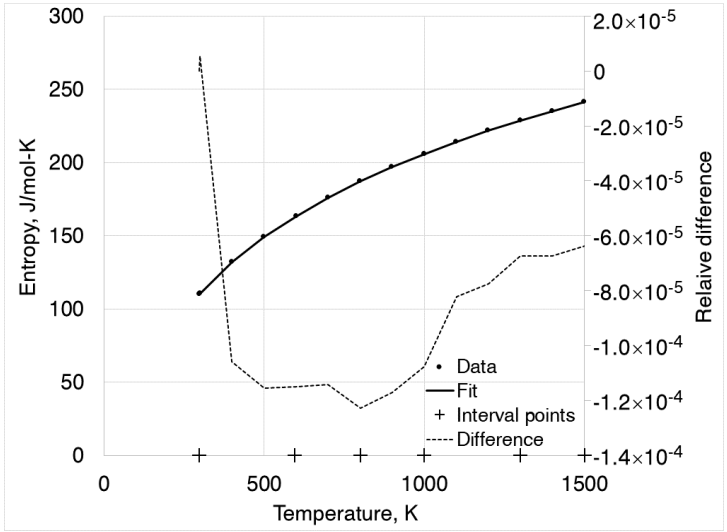


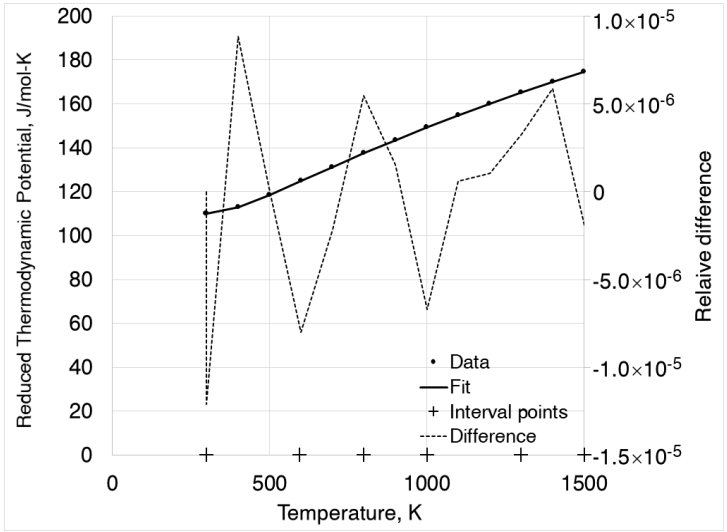
Zirconium Chloride ($ZrCl_2$) – Cl2Zr1(cr)

MW = 162.13 g/mol – $\Delta_f H^o$(298.15 K) = -430.952 kJ/mol – S^o(298.15 K) = 110.039 J/mol-K – p^o = 0.1 MPa



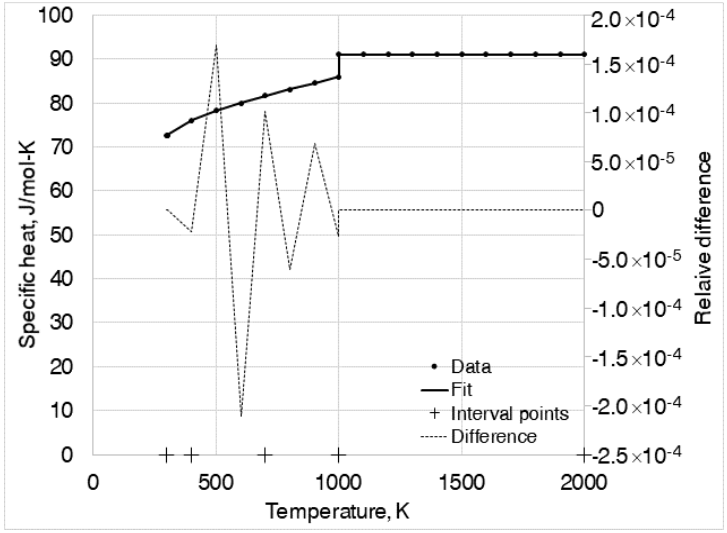


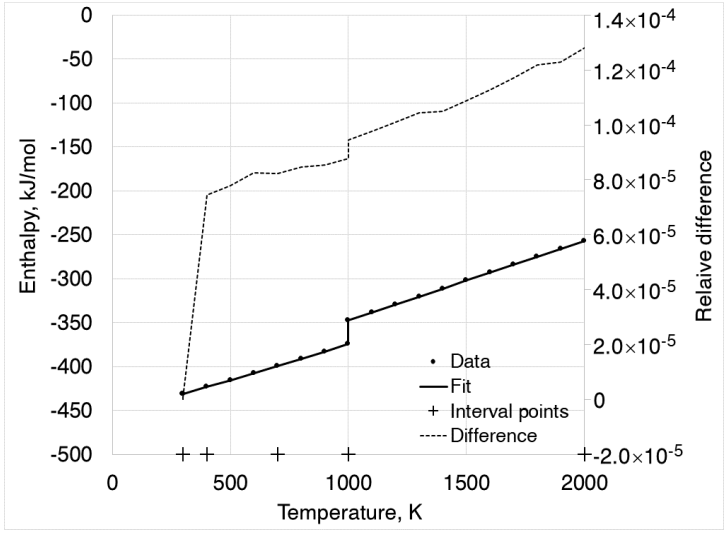


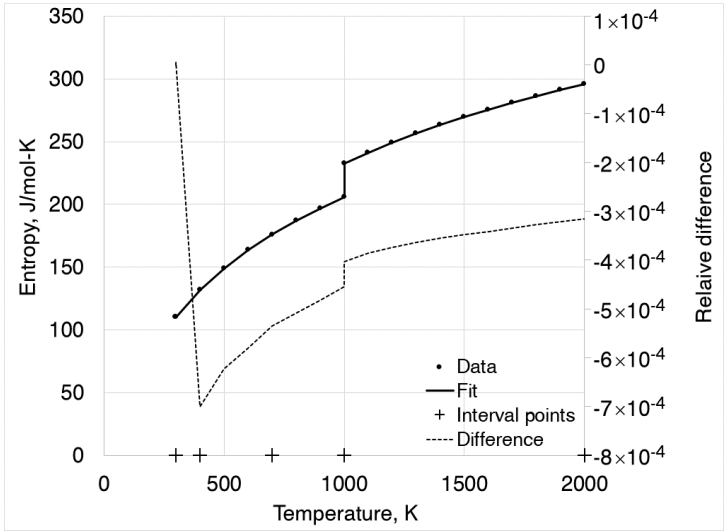


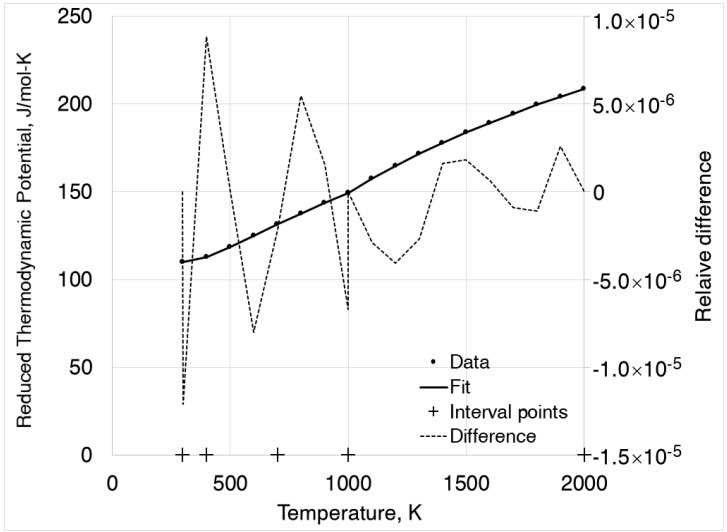
Zirconium Chloride ($ZrCl_2$) – Cl2Zr1(cr,l)

MW = 162.13 g/mol – $\Delta_f H^o$(298.15 K) = -430.952 kJ/mol – S^o(298.15 K) = 110.039 J/mol-K – p^o = 0.1 MPa



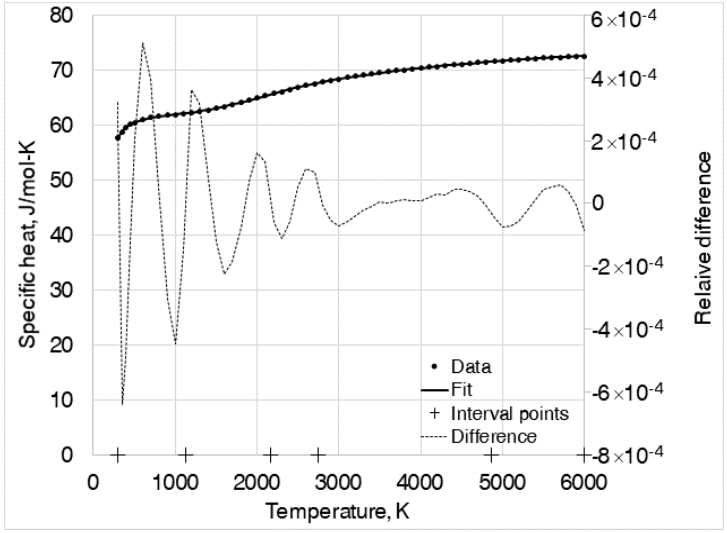


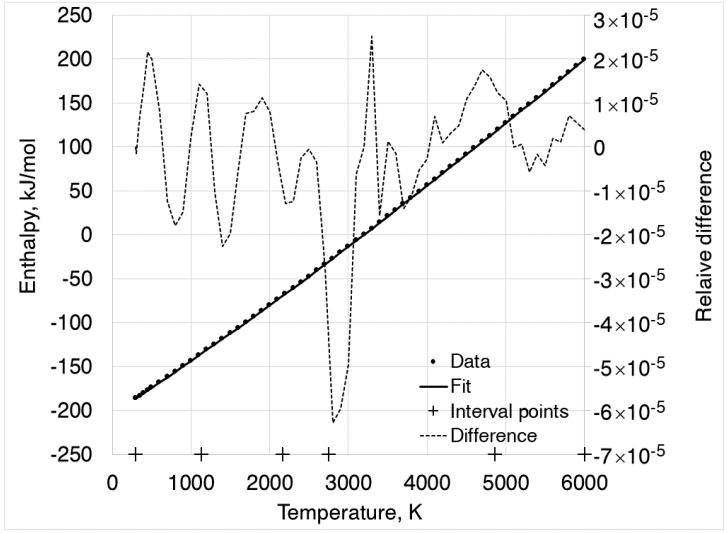


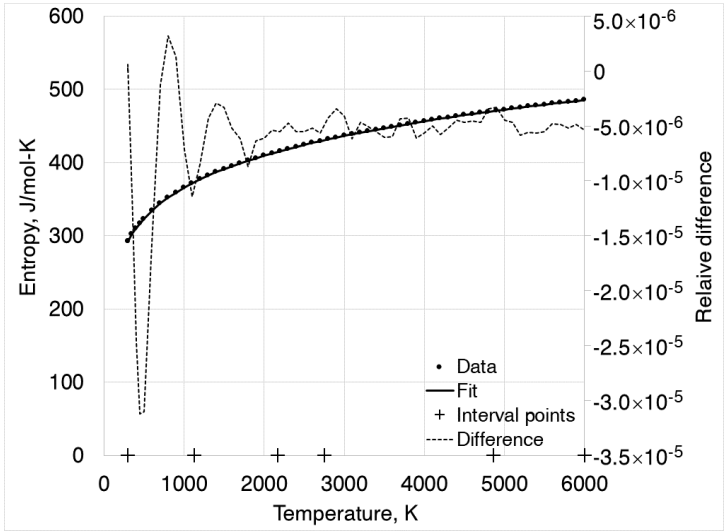


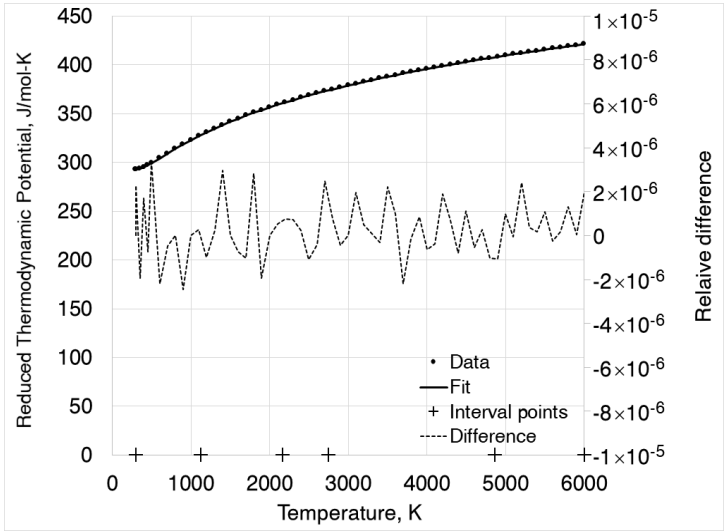
Zirconium Chloride ($ZrCl_2$) – Cl2Zr1(g)

MW = 162.13 g/mol – $\Delta_f H^o$(298.15 K) = -186.188 kJ/mol – S^o(298.15 K) = 292.558 J/mol-K – p^o = 0.1 MPa



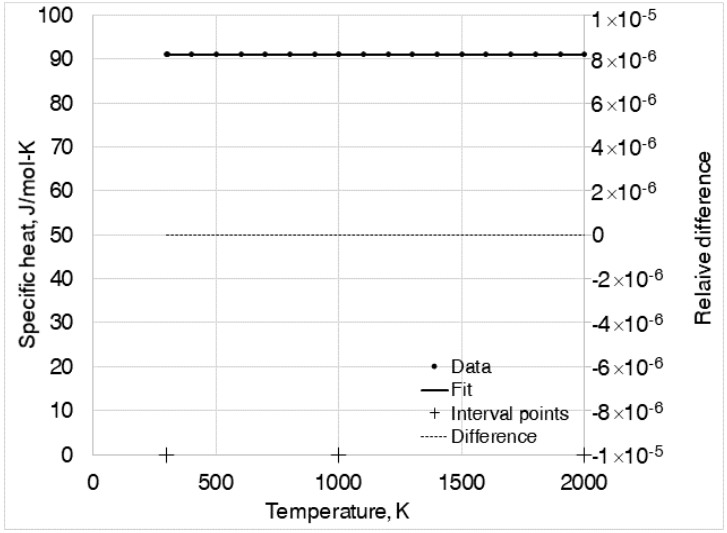


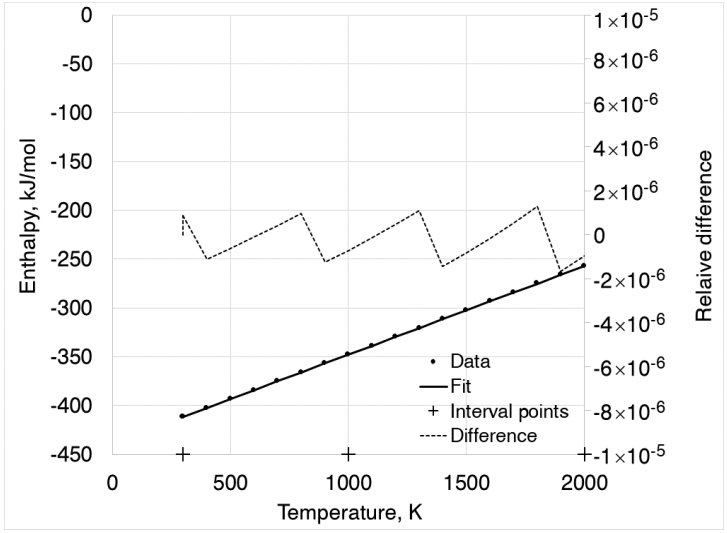


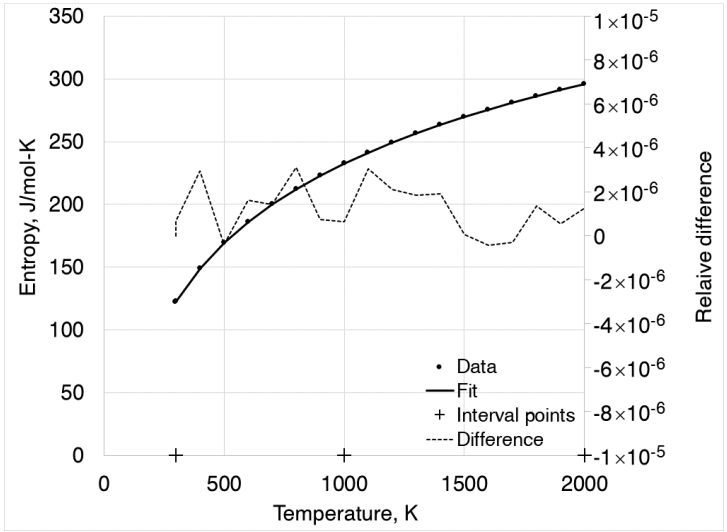


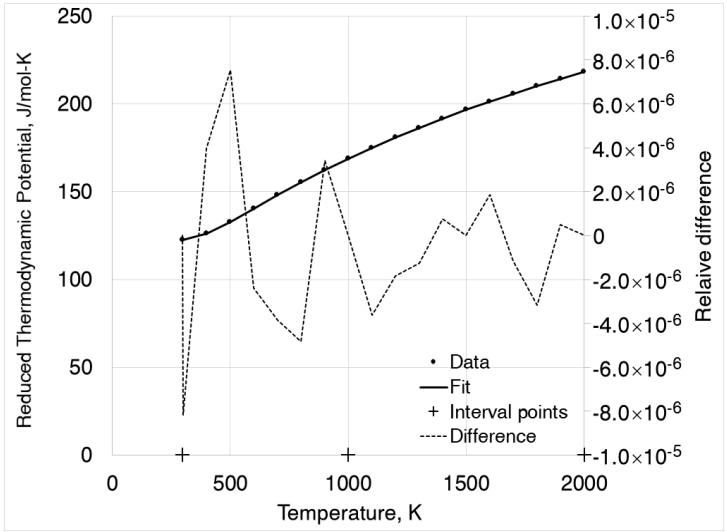
Zirconium Chloride ($ZrCl_2$) – Cl2Zr1(l)

MW = 162.13 g/mol – $\Delta_f H^o$(298.15 K) = -411.594 kJ/mol – S^o(298.15 K) = 122.493 J/mol-K – p^o = 0.1 MPa



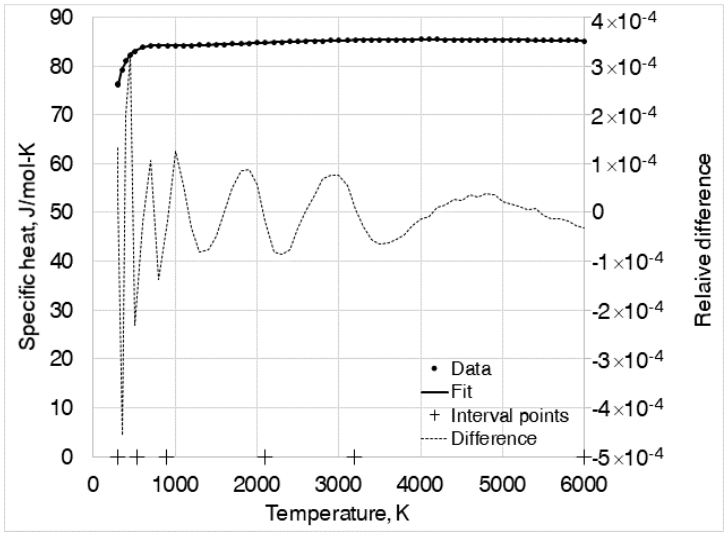


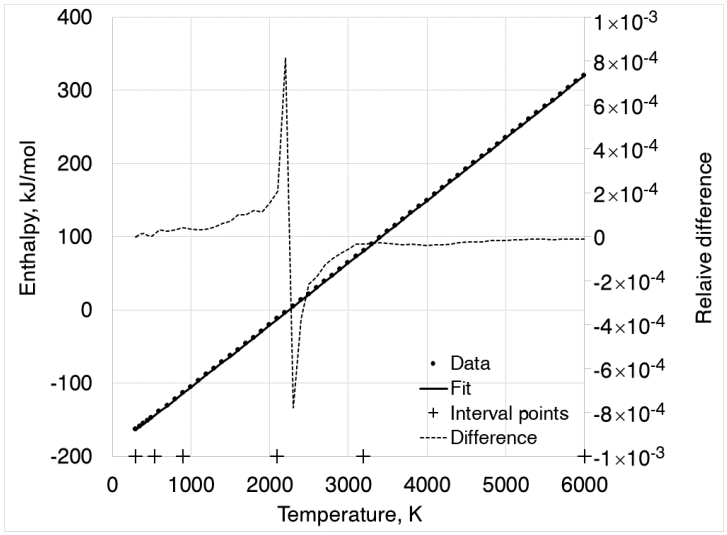


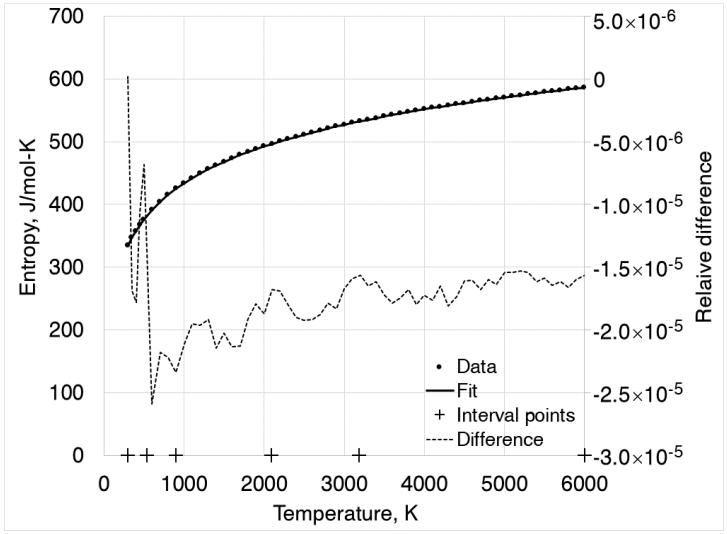


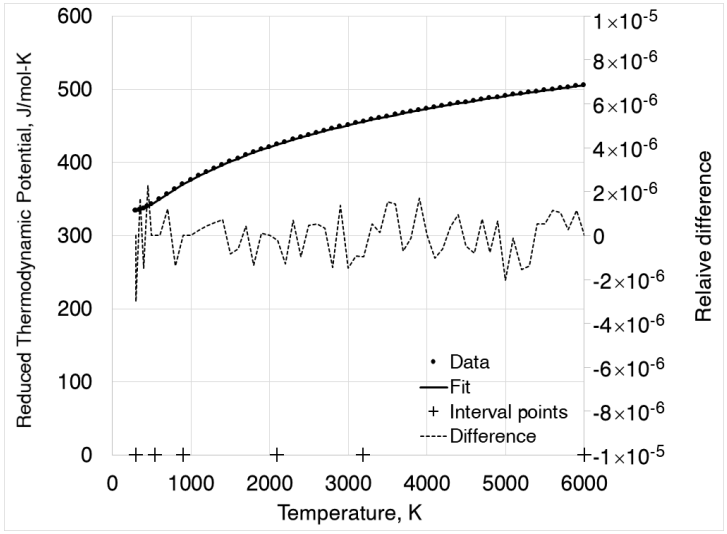
Cobalt Chloride ($CoCl_3$) – Cl3Co1(g)

MW = 165.29 g/mol – $\Delta_f H^o$(298.15 K) = -163.594 kJ/mol – S^o(298.15 K) = 334.209 J/mol-K – p^o = 0.1 MPa



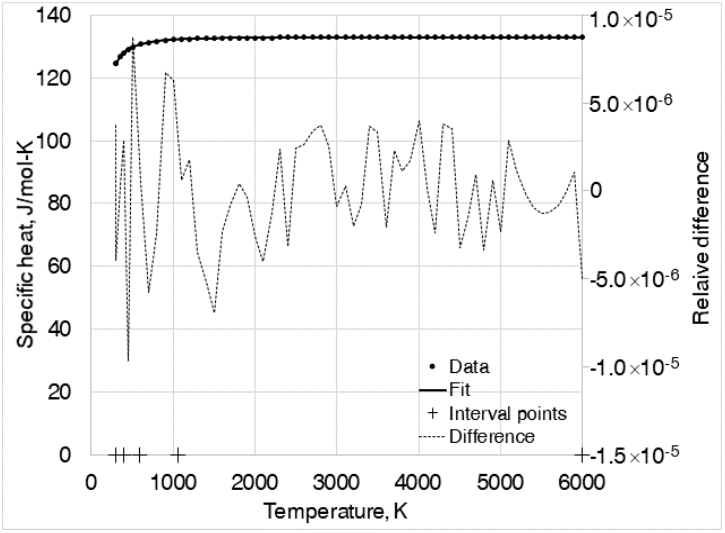


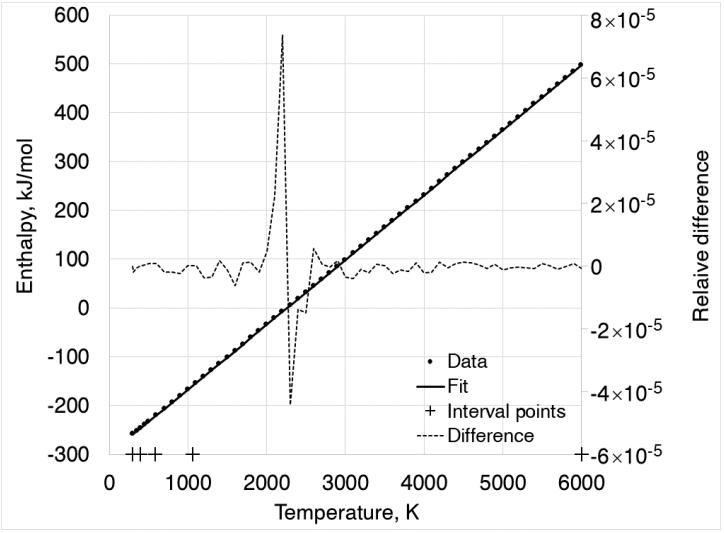


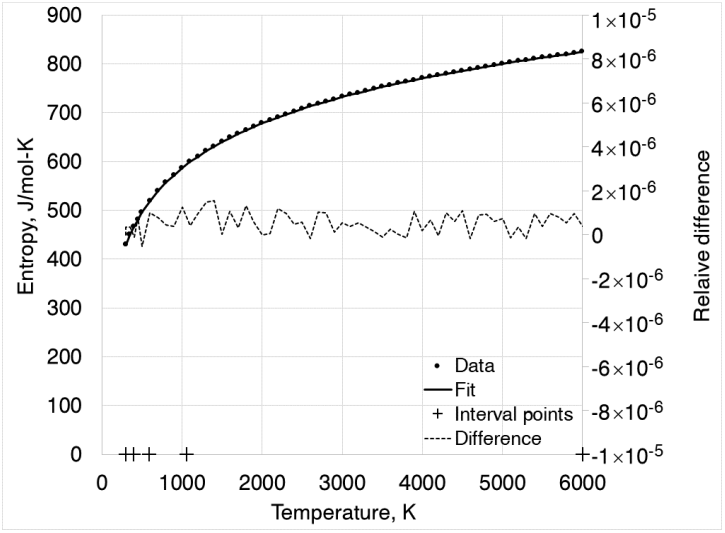


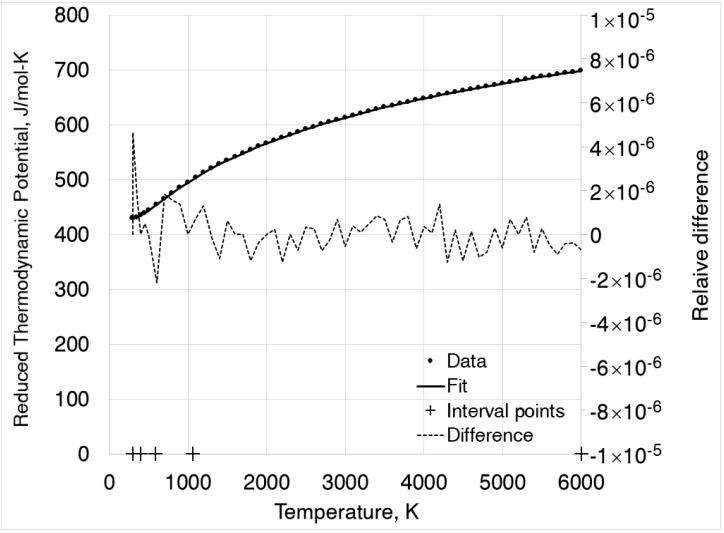
Copper Chloride ($(CuCl)_3$) – Cl3Cu3(g)

MW = 296.997 g/mol – $\Delta_f H^o$(298.15 K) = -258.571 kJ/mol – S^o(298.15 K) = 429.526 J/mol-K – p^o = 0.1 MPa



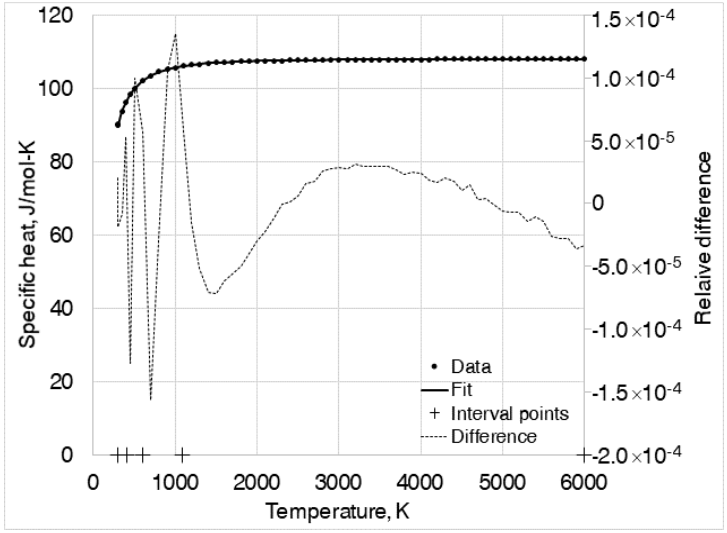


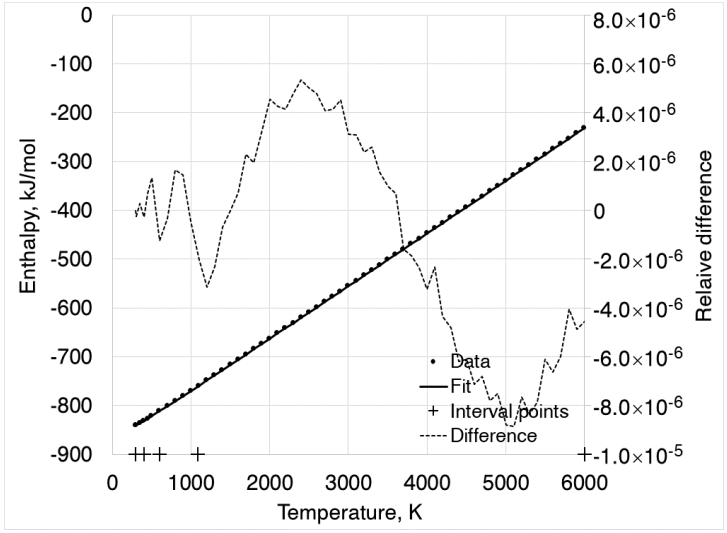


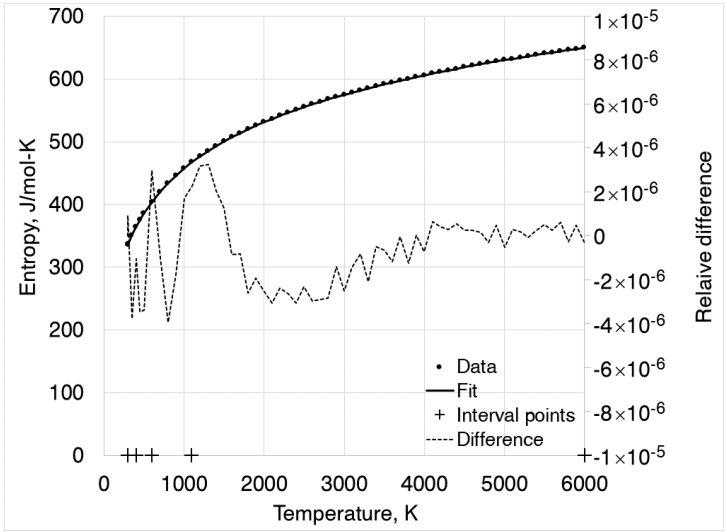


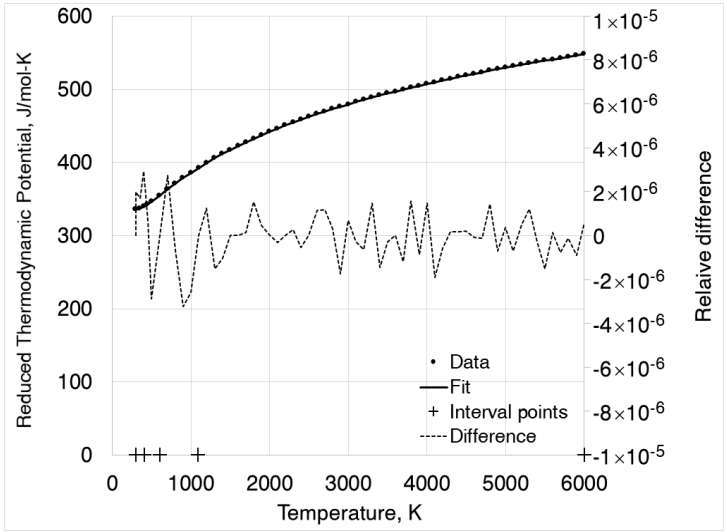
Trichlorofluorosilane ($SiCl_3F$) – Cl3F1Si1(g)

MW = 153.443 g/mol – $\Delta_f H^o$(298.15 K) = -840.984 kJ/mol – S^o(298.15 K) = 336.055 J/mol-K – p^o = 0.1 MPa



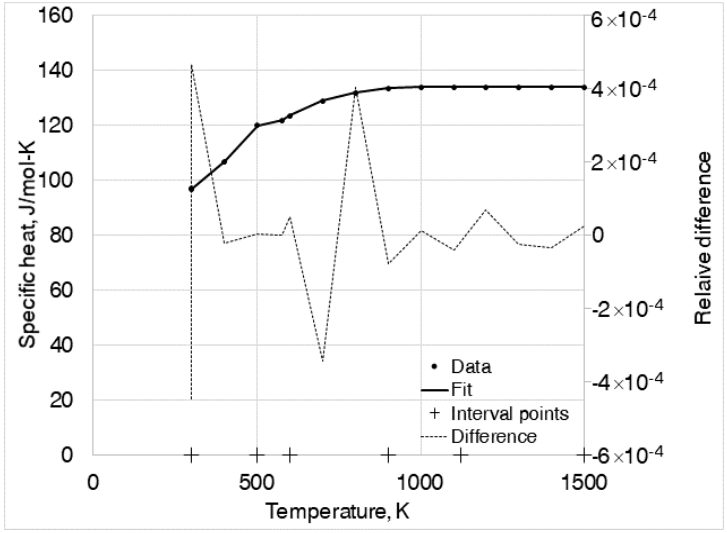


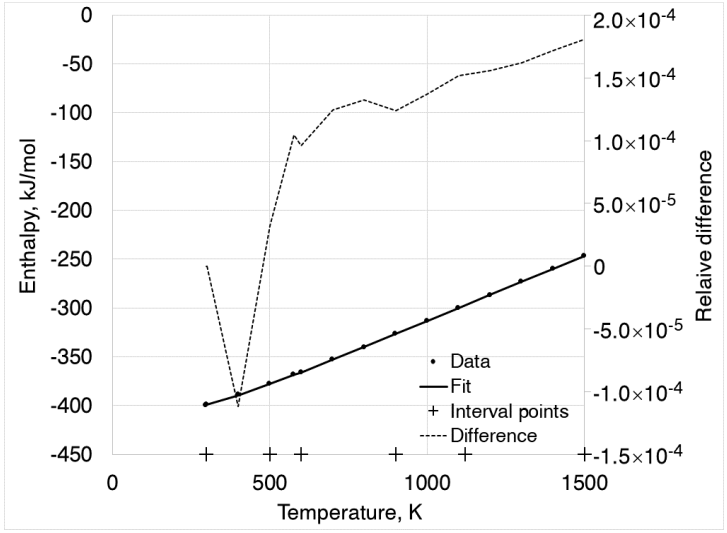


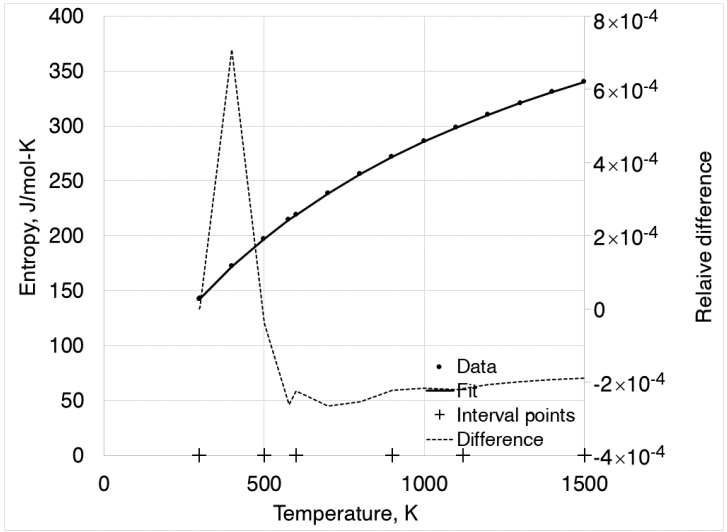


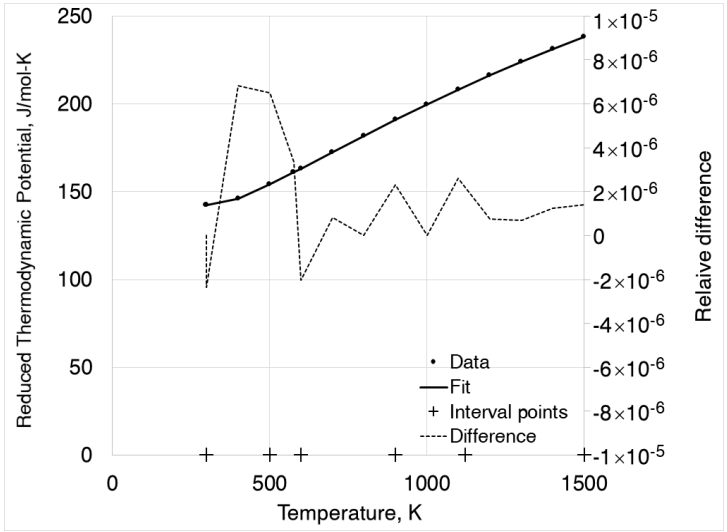
Iron Chloride ($FeCl_3$) – Cl3Fe1(cr)

MW = 162.206 g/mol – $\Delta_f H^o$(298.15 K) = -399.405 kJ/mol – S^o(298.15 K) = 142.336 J/mol-K – p^o = 0.1 MPa



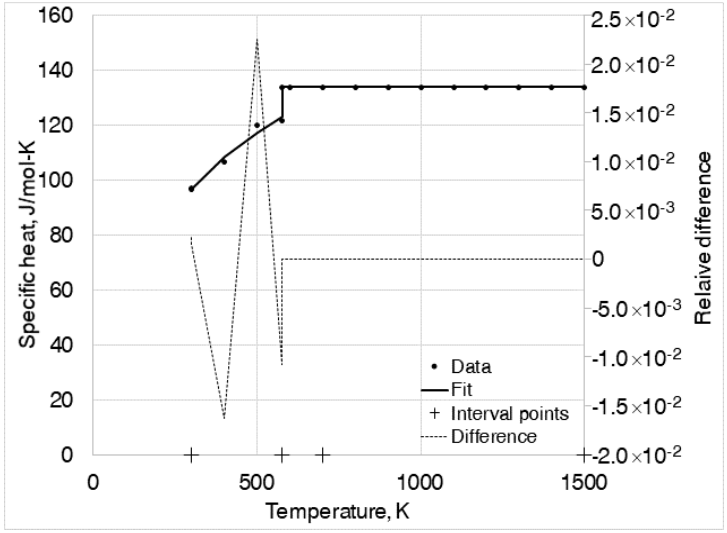


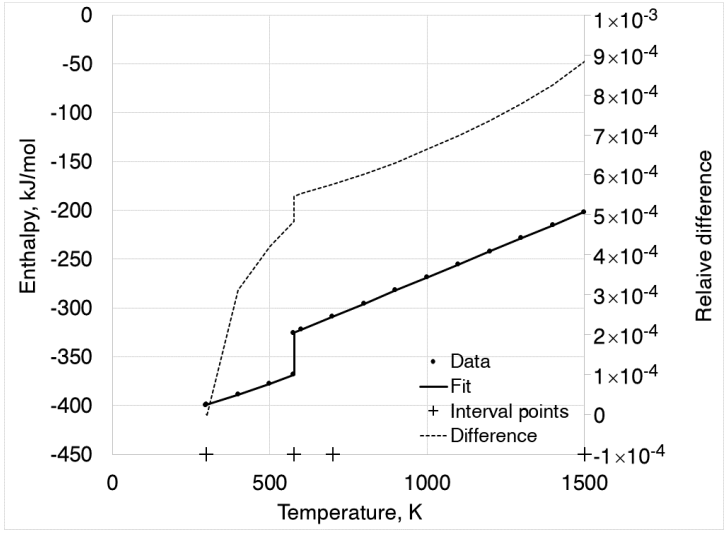


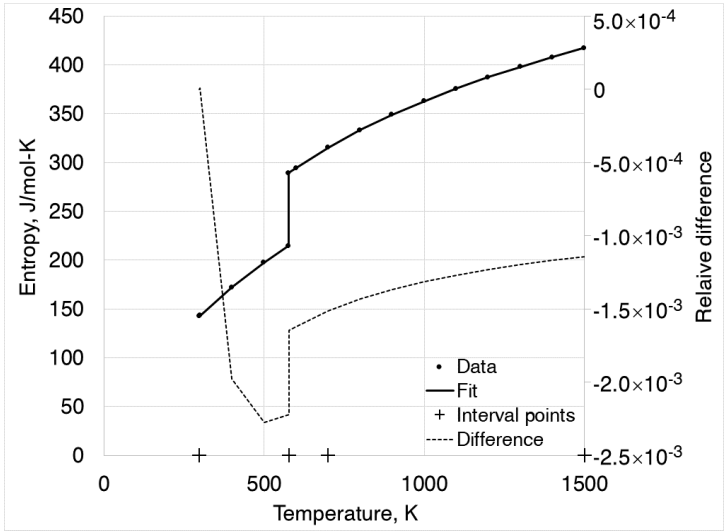


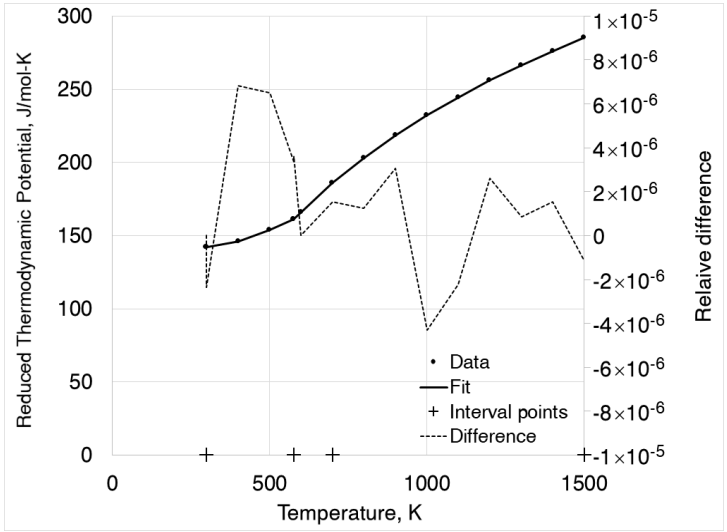
Iron Chloride ($FeCl_3$) – Cl3Fe1(cr,l)

MW = 162.206 g/mol – $\Delta_f H^o$(298.15 K) = -399.405 kJ/mol – S^o(298.15 K) = 142.336 J/mol-K – p^o = 0.1 MPa



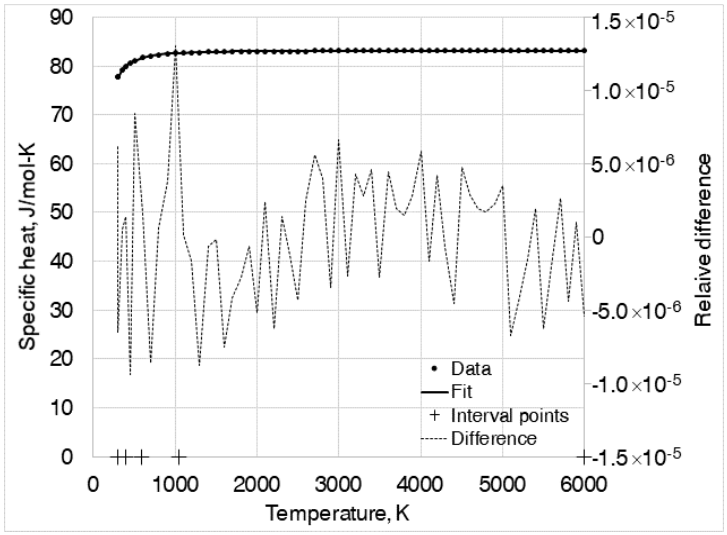


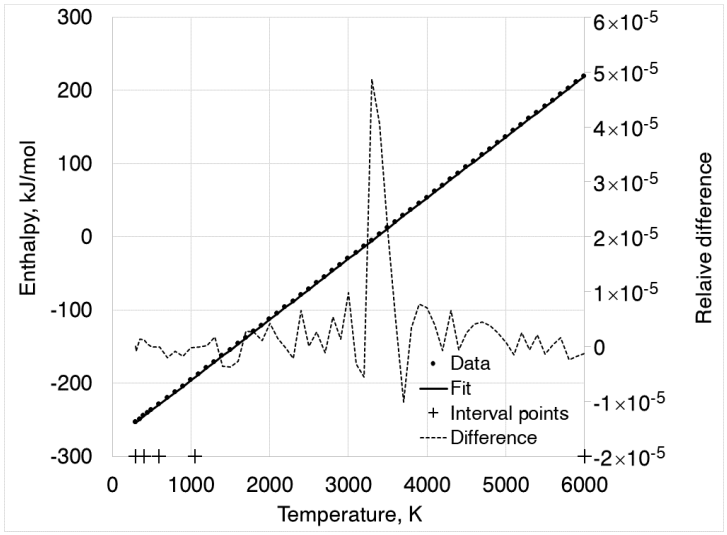


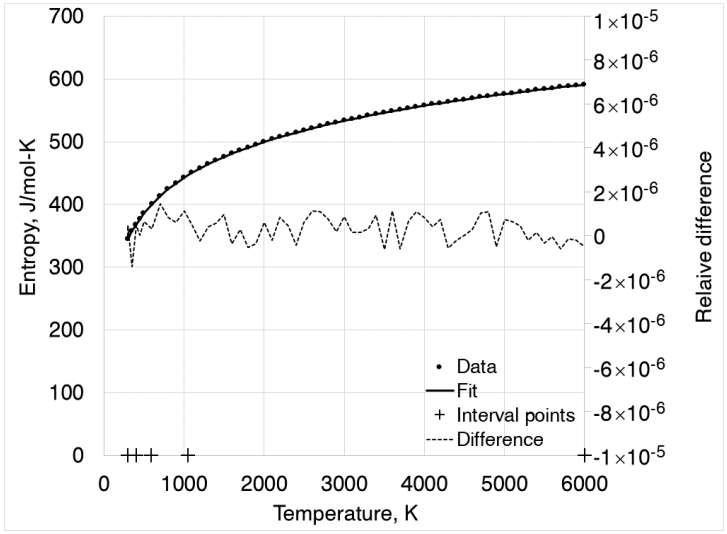


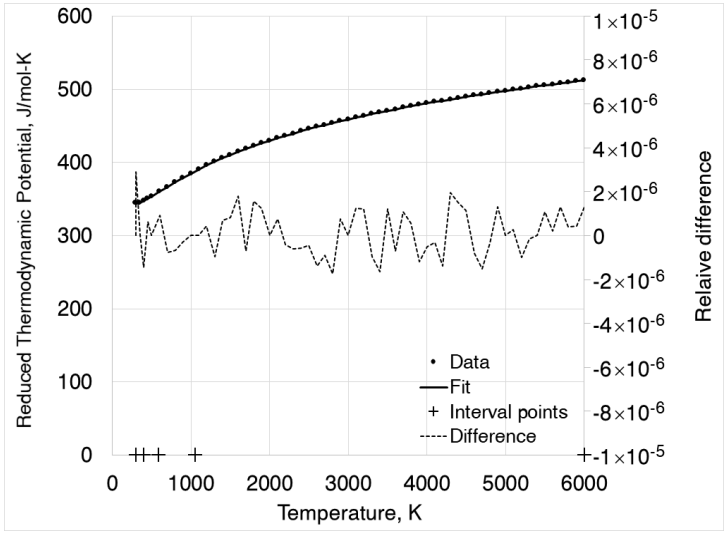
Iron Chloride ($FeCl_3$) – Cl3Fe1(g)

MW = 162.206 g/mol – $\Delta_f H^o$(298.15 K) = -253.132 kJ/mol – S^o(298.15 K) = 344.206 J/mol-K – p^o = 0.1 MPa



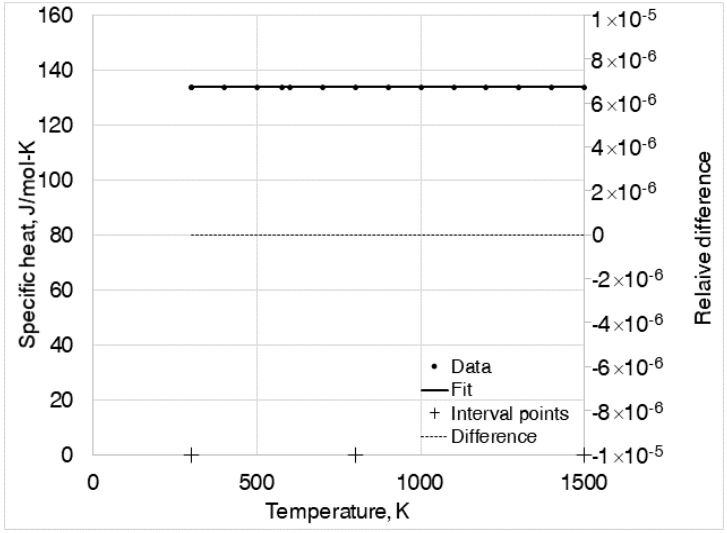


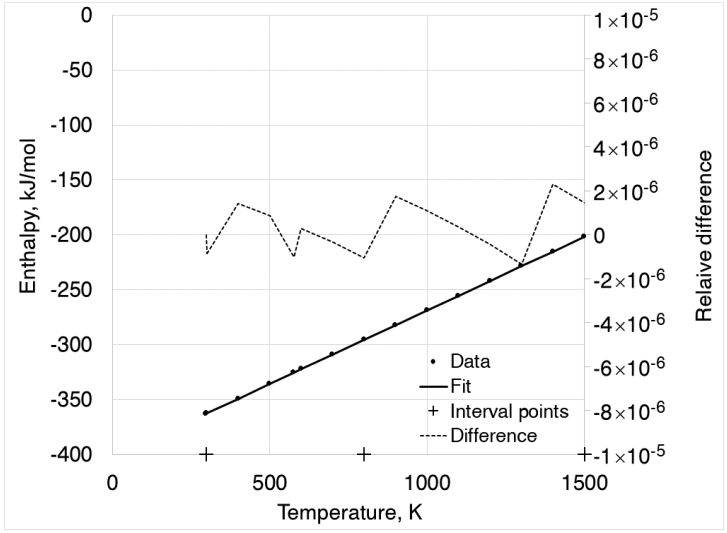


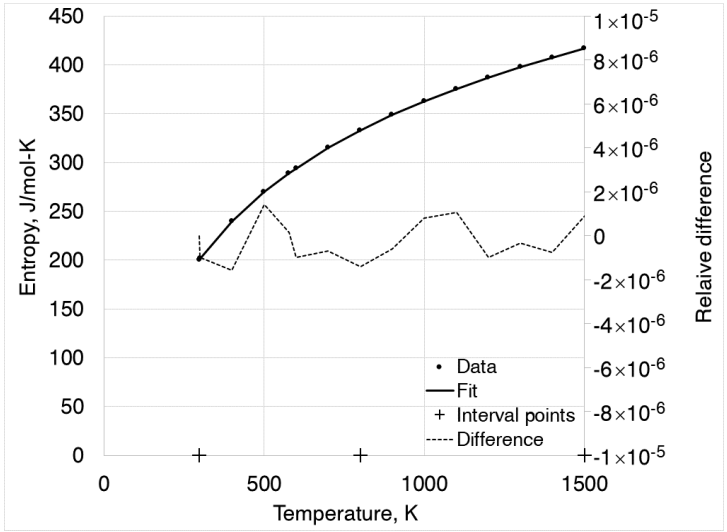


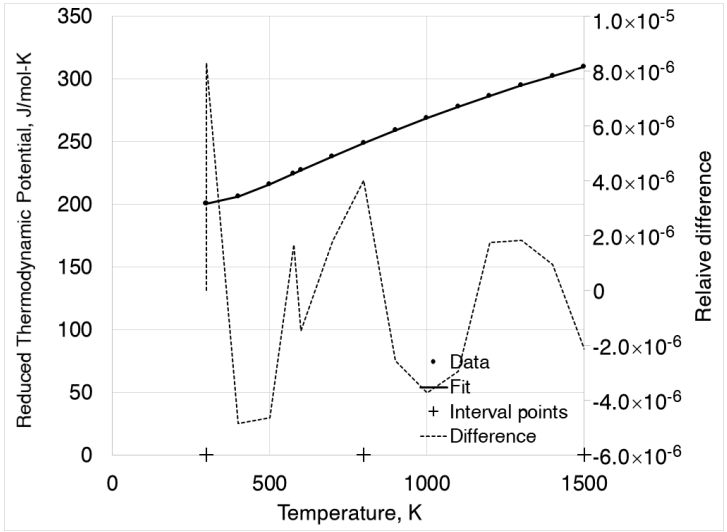
Iron Chloride ($FeCl_3$) – Cl3Fe1(l)

MW = 162.206 g/mol – $\Delta_f H^o$(298.15 K) = -362.815 kJ/mol – S^o(298.15 K) = 200.657 J/mol-K – p^o = 0.1 MPa



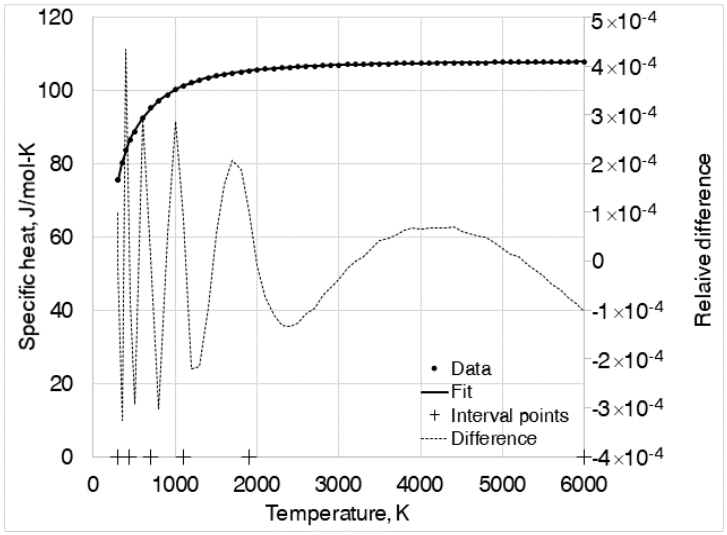


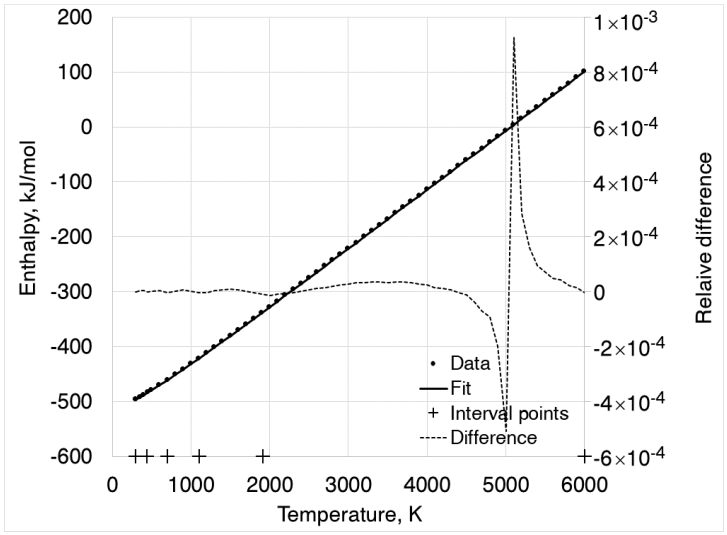


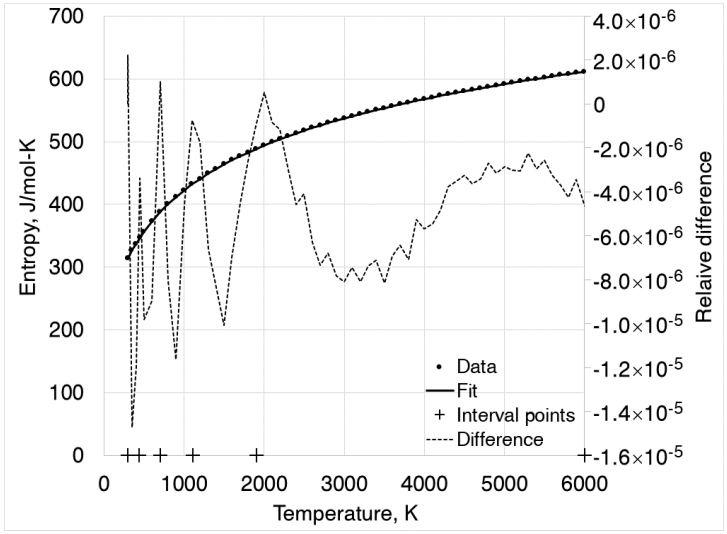


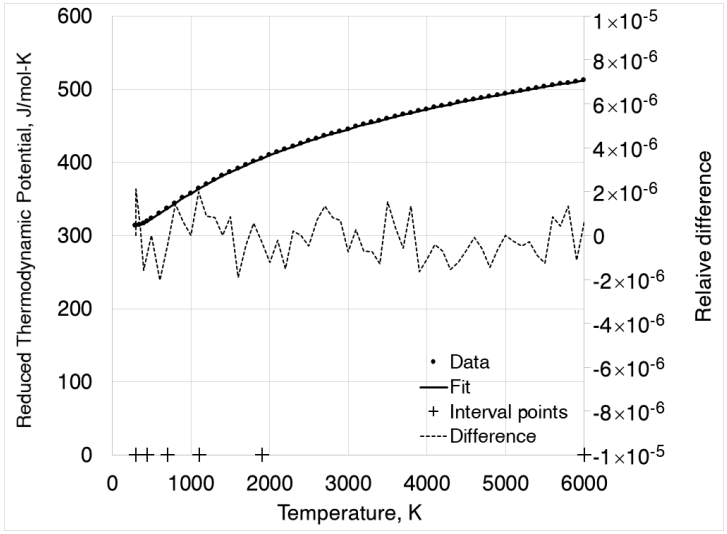
Trichlorosilane ($SiHCl_3$) – Cl3H1Si1(g)

MW = 135.452 g/mol – $\Delta_f H^o$(298.15 K) = -496.222 kJ/mol – S^o(298.15 K) = 313.717 J/mol-K – p^o = 0.1 MPa



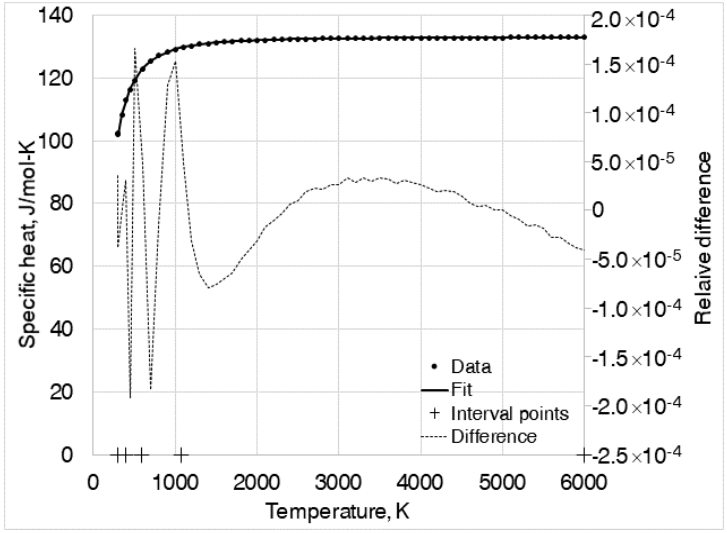


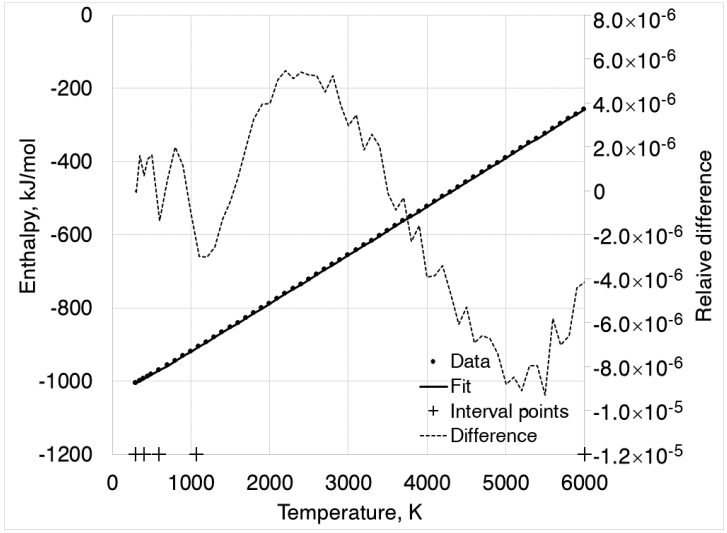


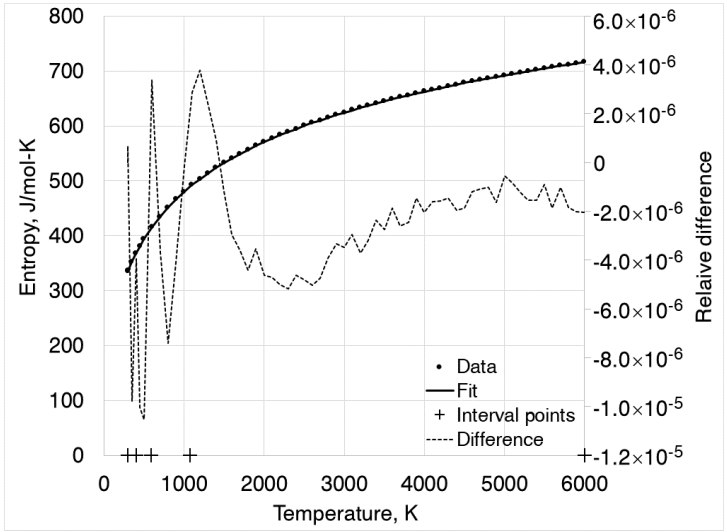


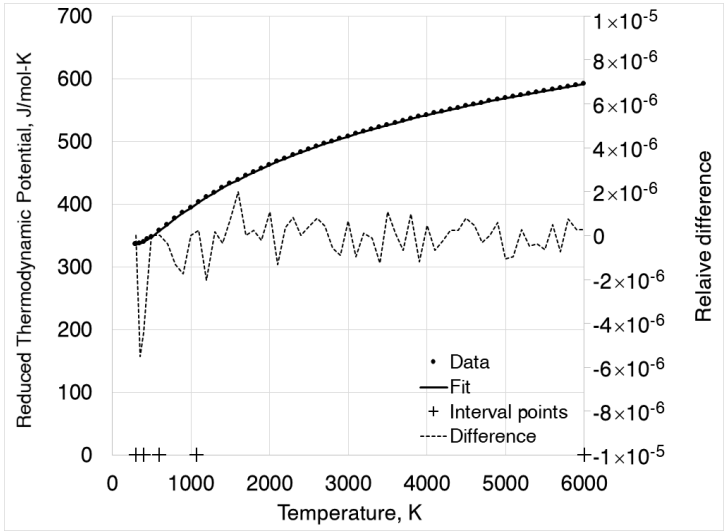
Lithium Chloride ($(LiCl)_3$) – Cl3Li3(g)

MW = 127.182 g/mol – $\Delta_f H^o$(298.15 K) = -1004.721 kJ/mol – S^o(298.15 K) = 335.747 J/mol-K – p^o = 0.1 MPa



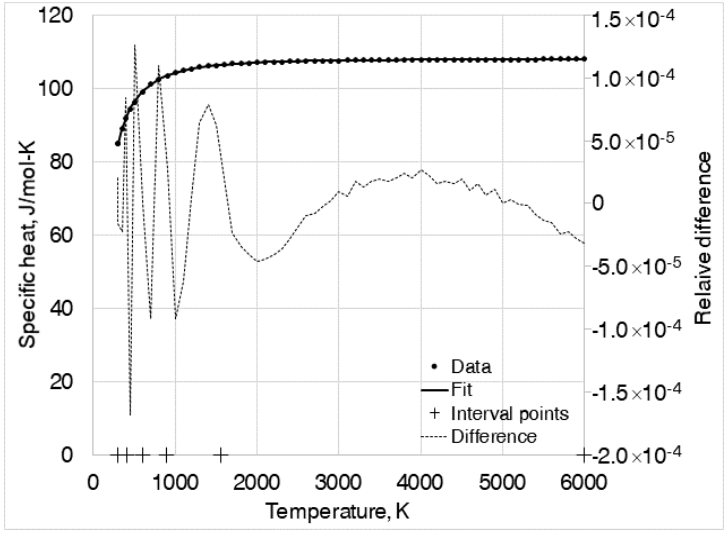


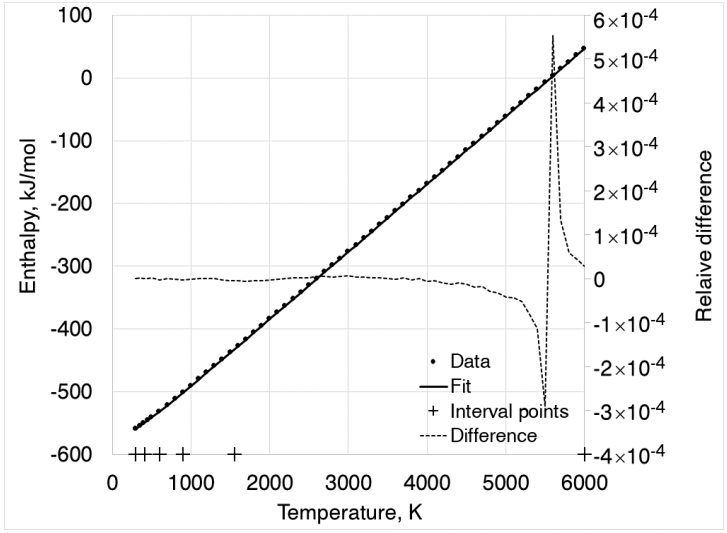


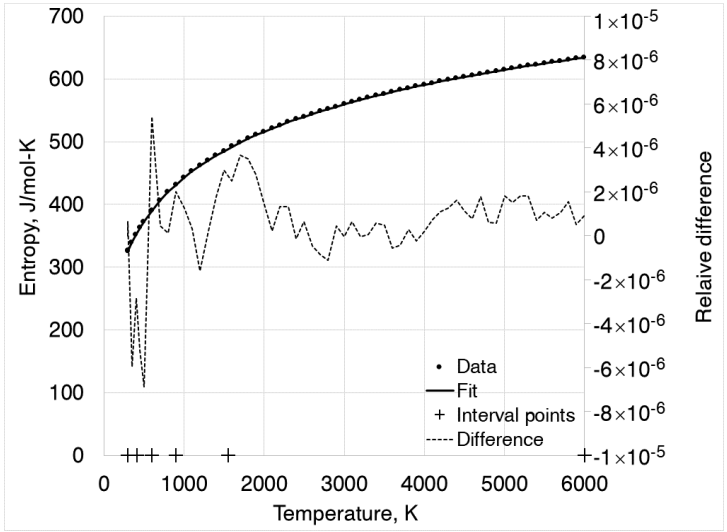


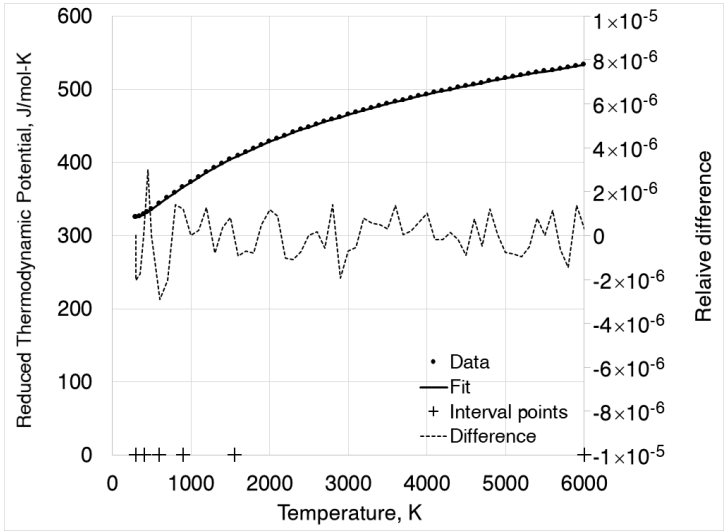
Phosphoryl Chloride ($OPCl_3$) – Cl3O1P1(g)

MW = 153.332 g/mol – $\Delta_f H^o$(298.15 K) = -559.819 kJ/mol – S^o(298.15 K) = 325.454 J/mol-K – p^o = 0.1 MPa



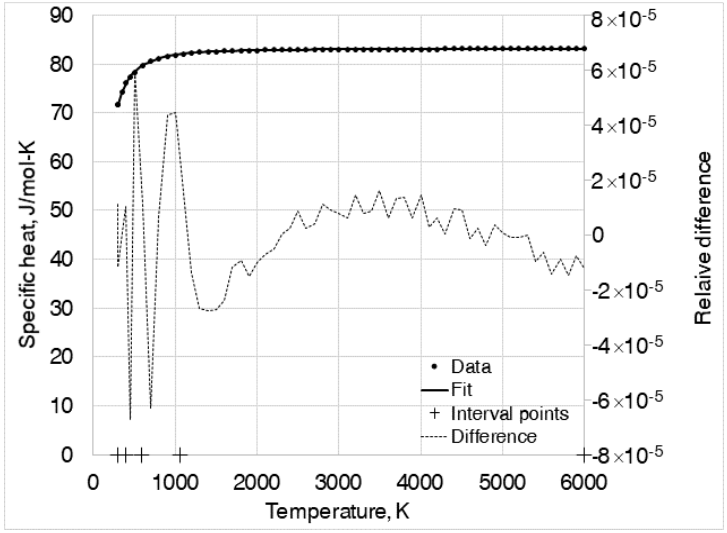


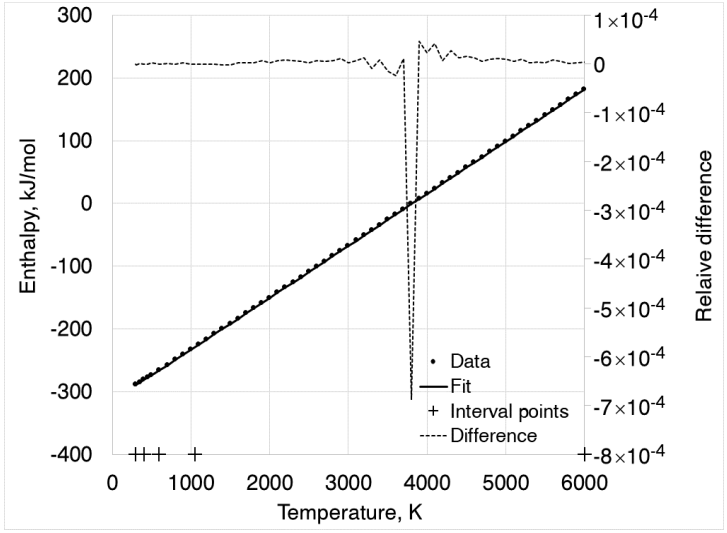


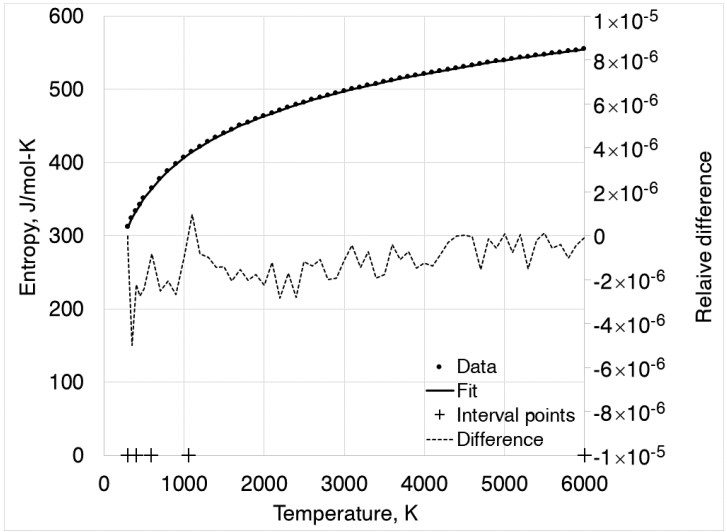


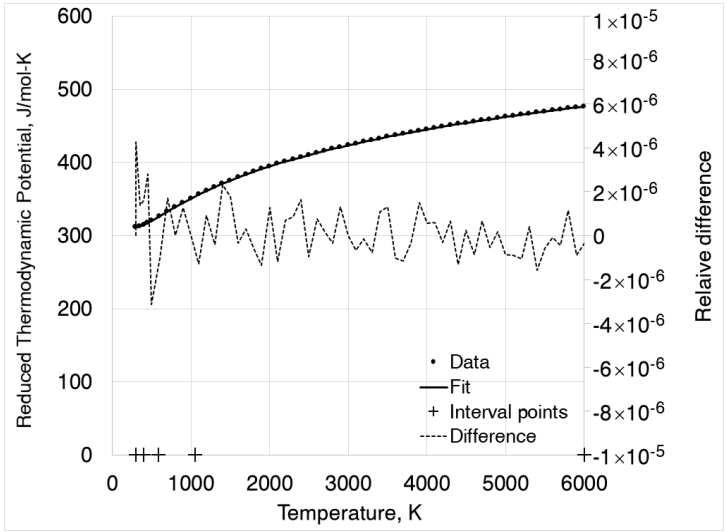
Phosphorus Chloride (PCl_3) – Cl3P1(g)

MW = 137.333 g/mol – $\Delta_f H^o$(298.15 K) = -288.696 kJ/mol – S^o(298.15 K) = 311.682 J/mol-K – p^o = 0.1 MPa



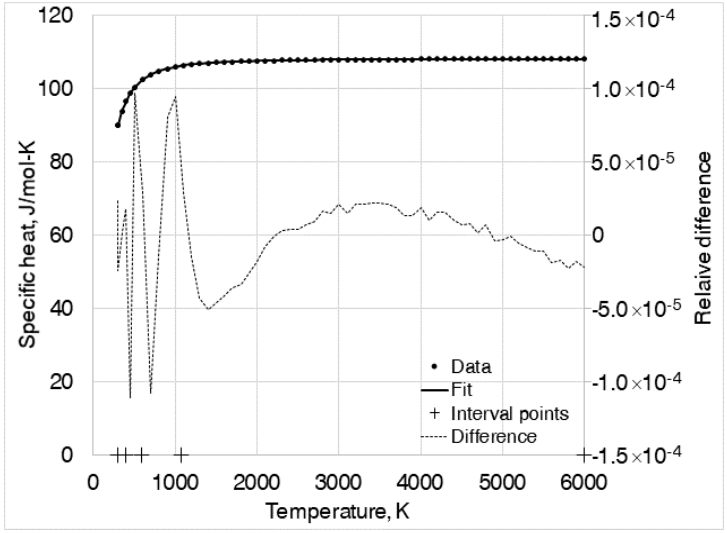


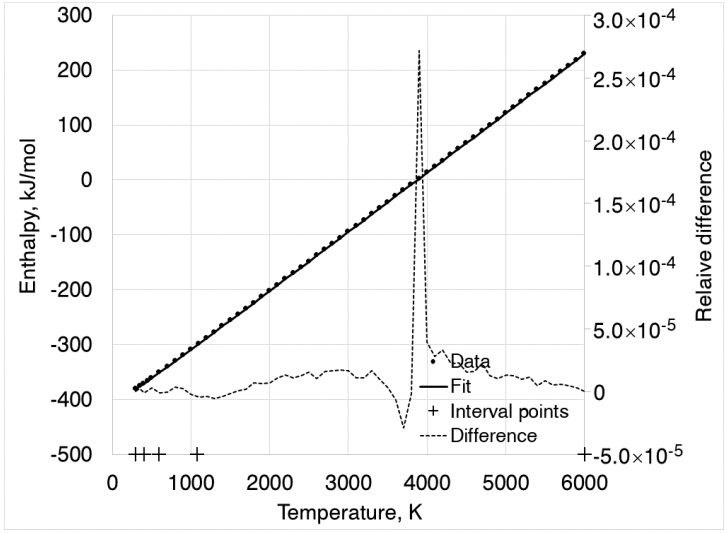


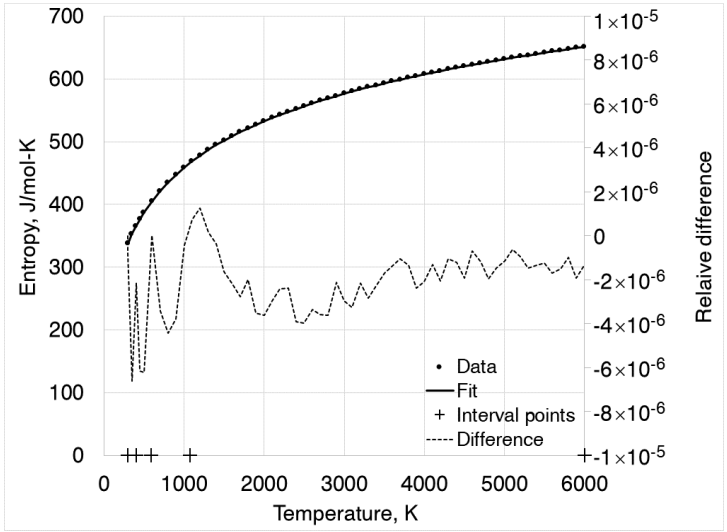


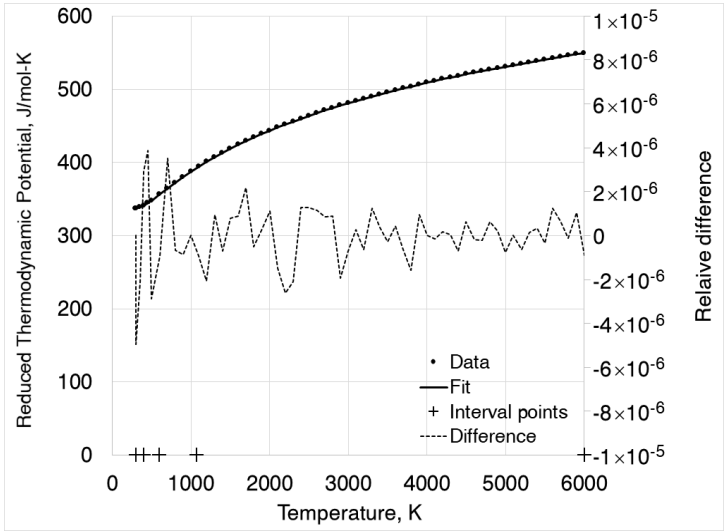
Thiophosphoryl Chloride ($SPCl_3$) – Cl3P1S1(g)

MW = 169.39 g/mol – $\Delta_f H^o$(298.15 K) = -380.744 kJ/mol – S^o(298.15 K) = 337.347 J/mol-K – p^o = 0.1 MPa



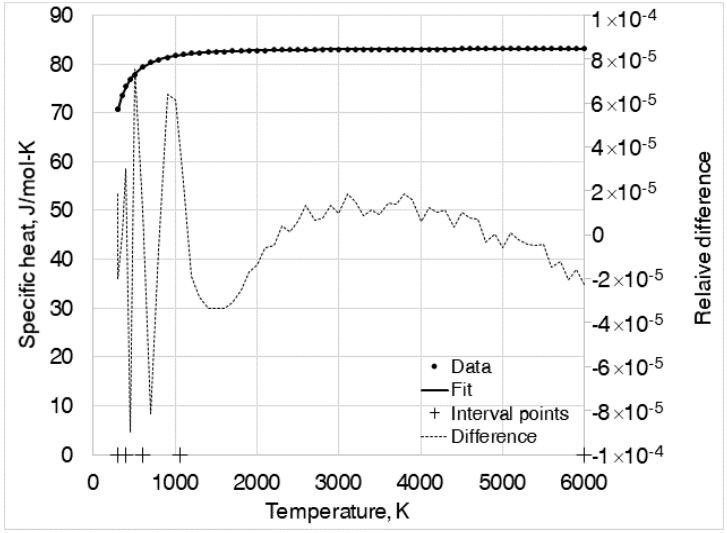


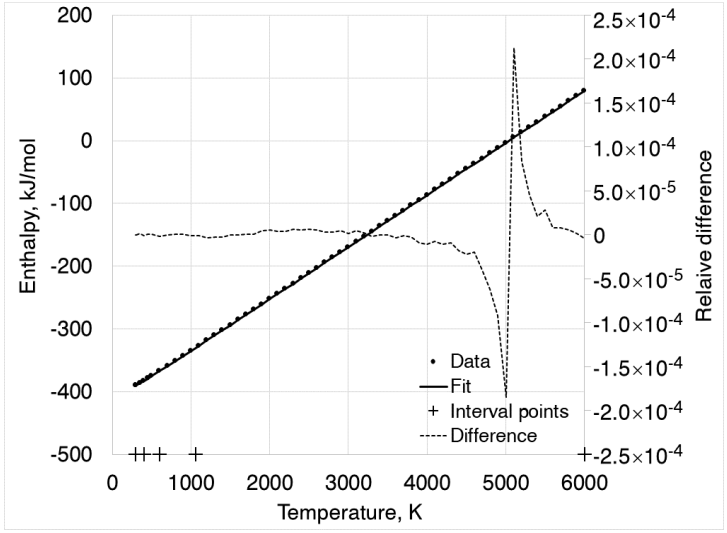


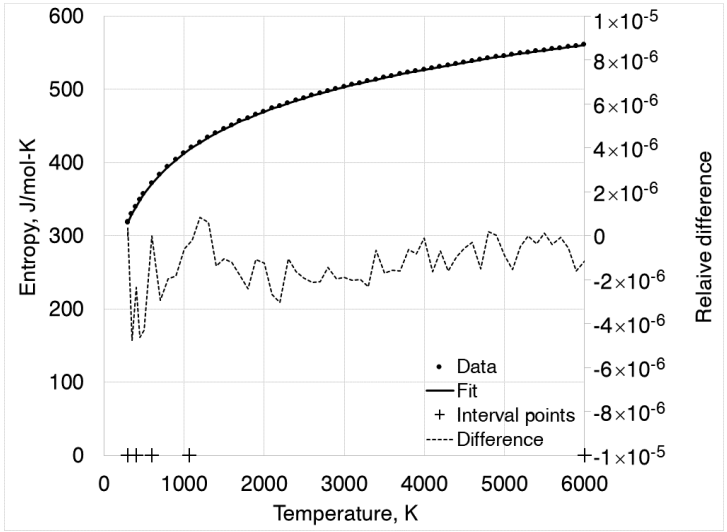


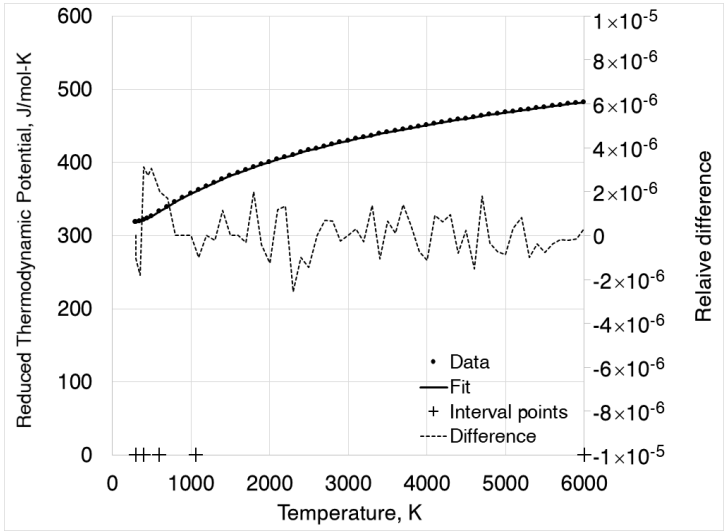
Trichlorosilyl ($SiCl_3$) – Cl3Si1(g)

MW = 134.444 g/mol – $\Delta_f H^o$(298.15 K) = -390.367 kJ/mol – S^o(298.15 K) = 318.189 J/mol-K – p^o = 0.1 MPa



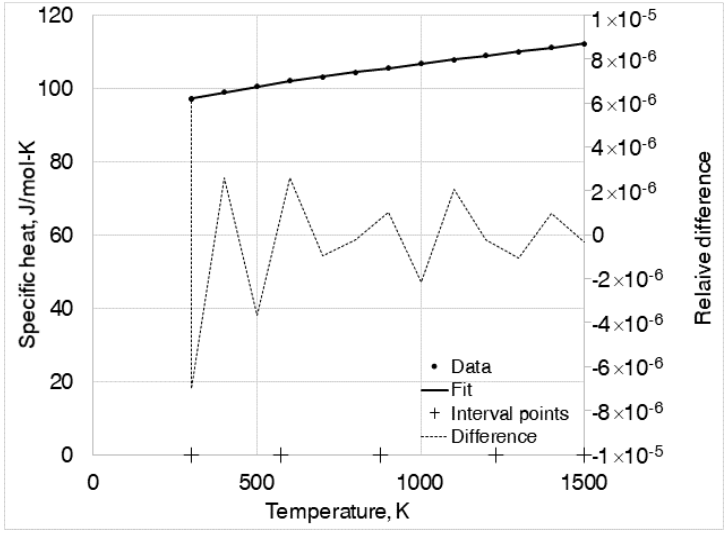


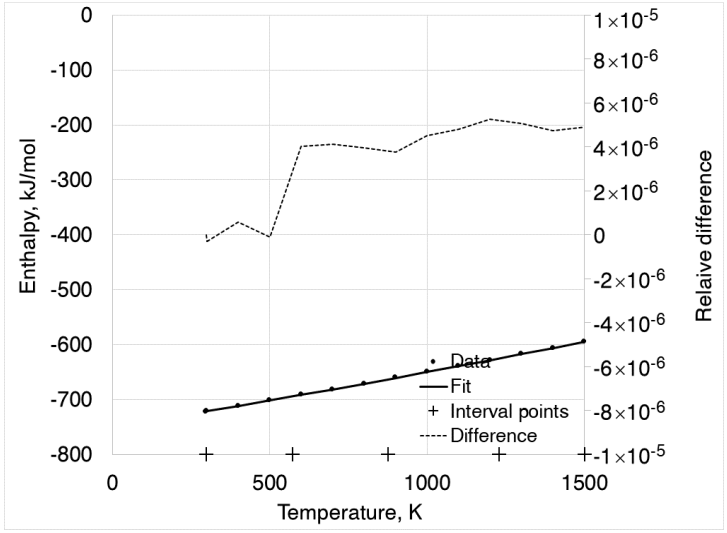


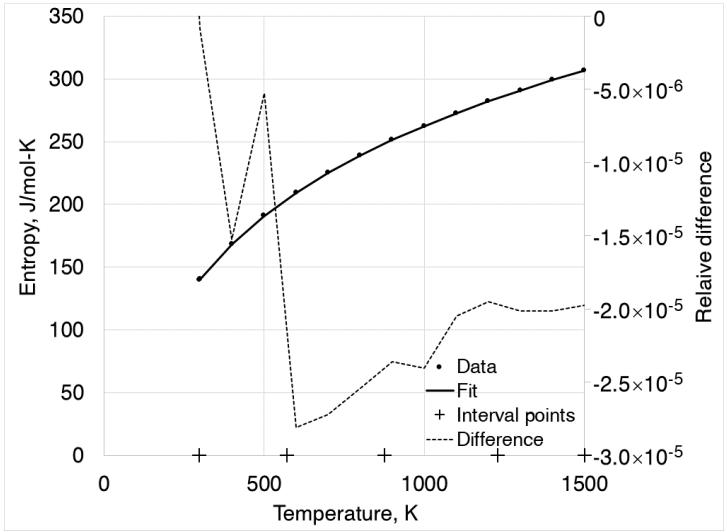


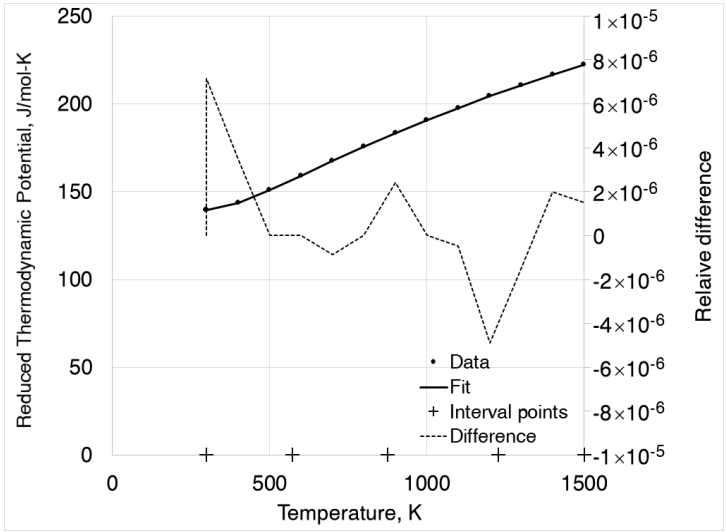
Titanium Chloride ($TiCl_3$) – Cl3Ti1(cr)

MW = 154.24 g/mol – $\Delta_f H^o$(298.15 K) = -721.74 kJ/mol – S^o(298.15 K) = 139.75 J/mol-K – p^o = 0.1 MPa



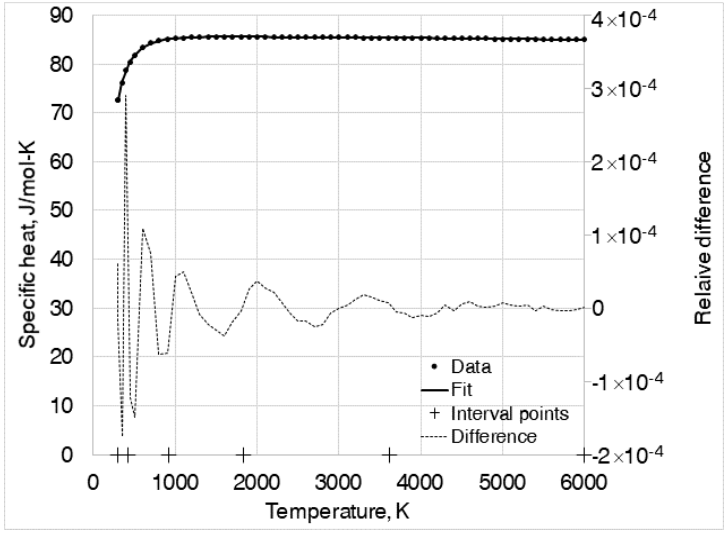


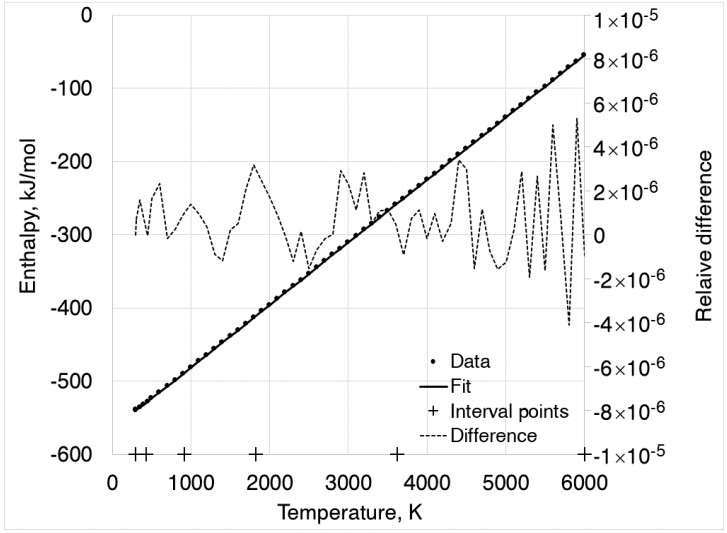


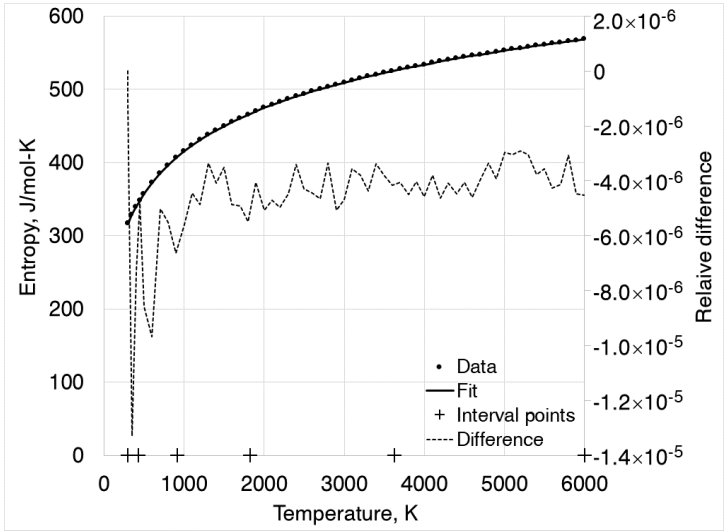


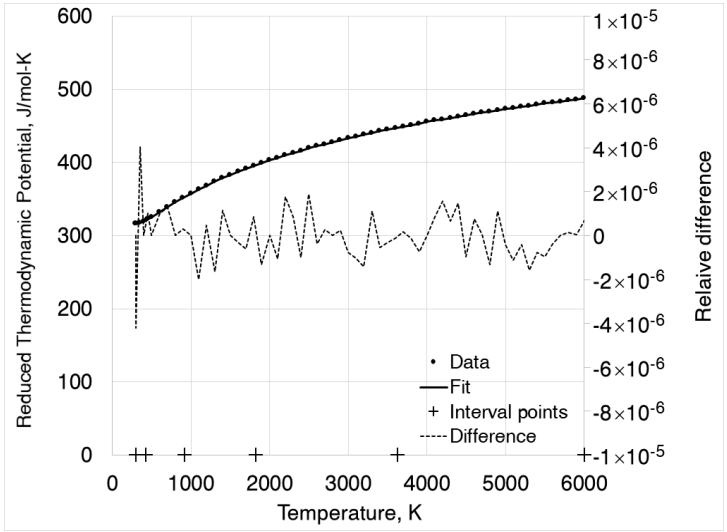
Titanium Chloride ($TiCl_3$) – Cl3Ti1(g)

MW = 154.24 g/mol – $\Delta_f H^o$(298.15 K) = -539.318 kJ/mol – S^o(298.15 K) = 316.895 J/mol-K – p^o = 0.1 MPa



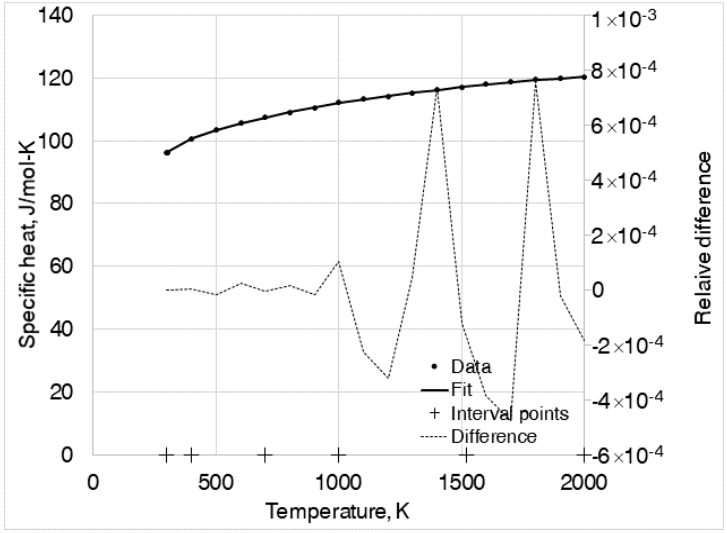


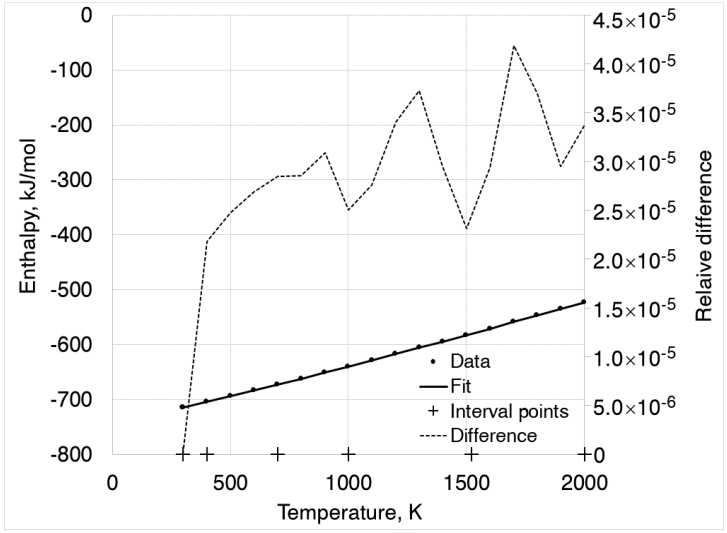


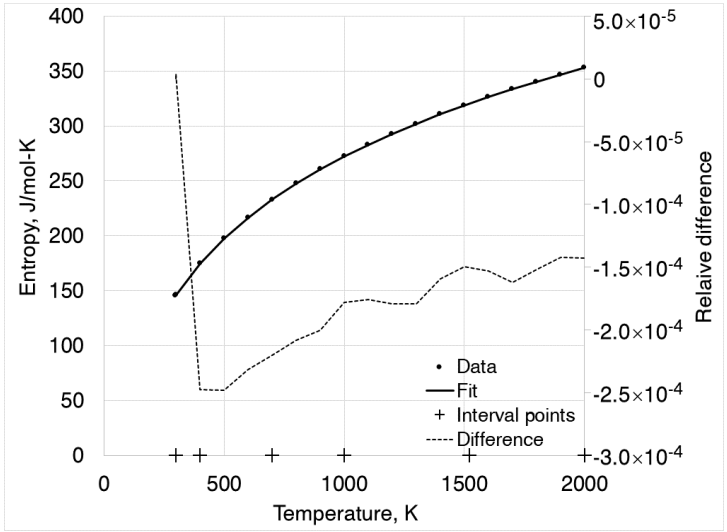


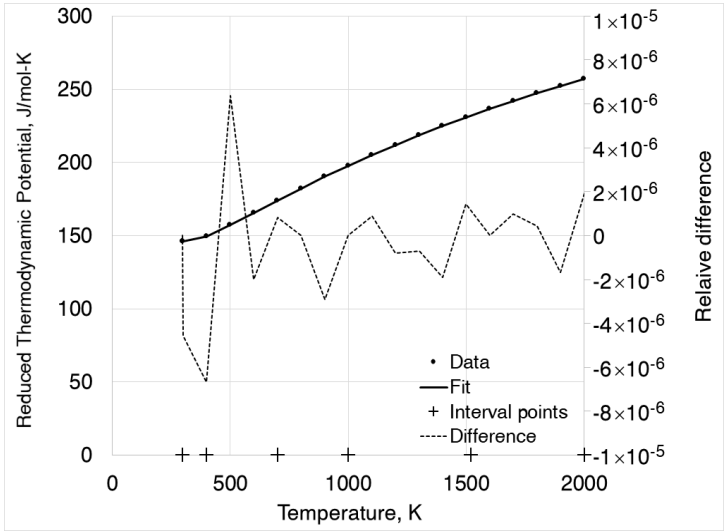
Zirconium Chloride ($ZrCl_3$) – Cl3Zr1(cr)

MW = 197.58 g/mol – $\Delta_f H^o$(298.15 K) = -714.209 kJ/mol – S^o(298.15 K) = 145.812 J/mol-K – p^o = 0.1 MPa



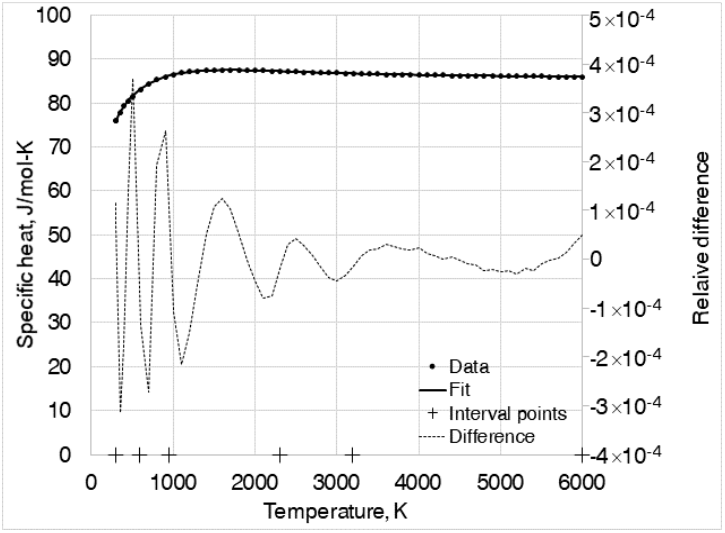


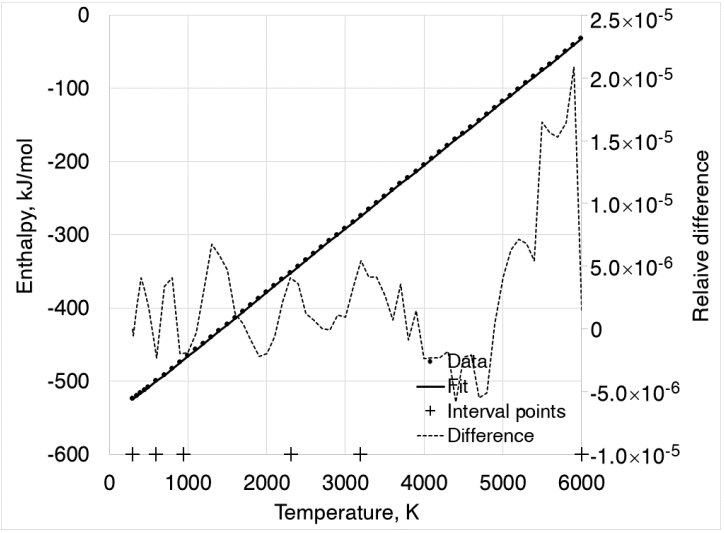


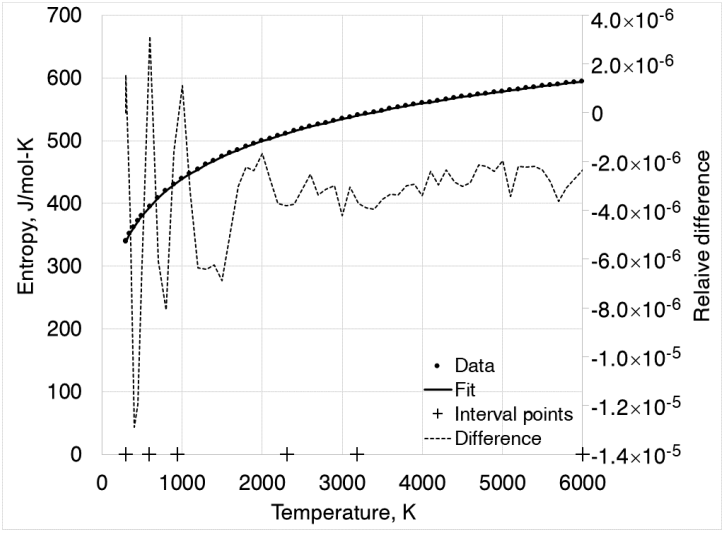


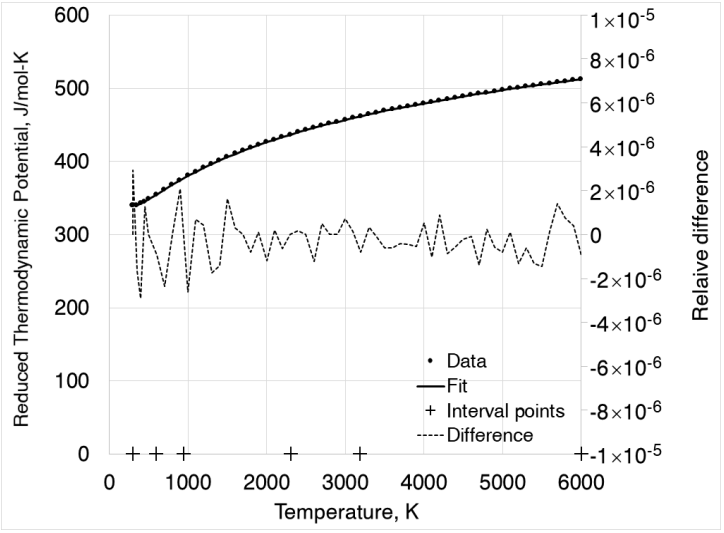
Zirconium Chloride ($ZrCl_3$) – Cl3Zr1(g)

MW = 197.58 g/mol – $\Delta_f H^o$(298.15 K) = -524.255 kJ/mol – S^o(298.15 K) = 339.34 J/mol-K – p^o = 0.1 MPa



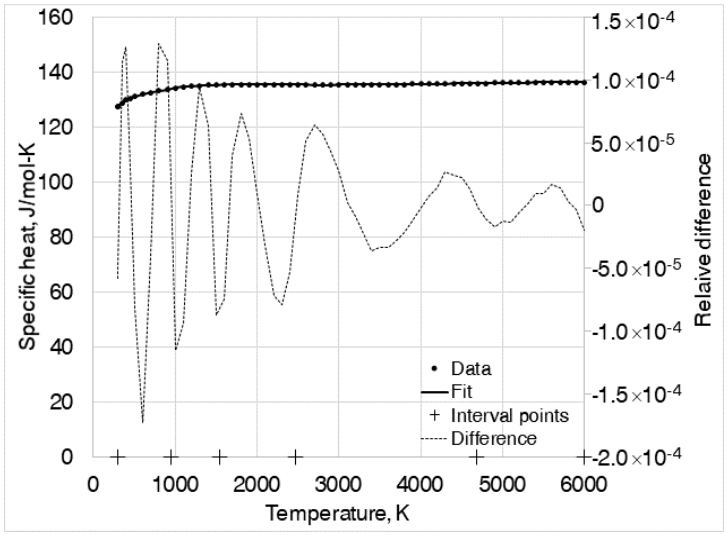


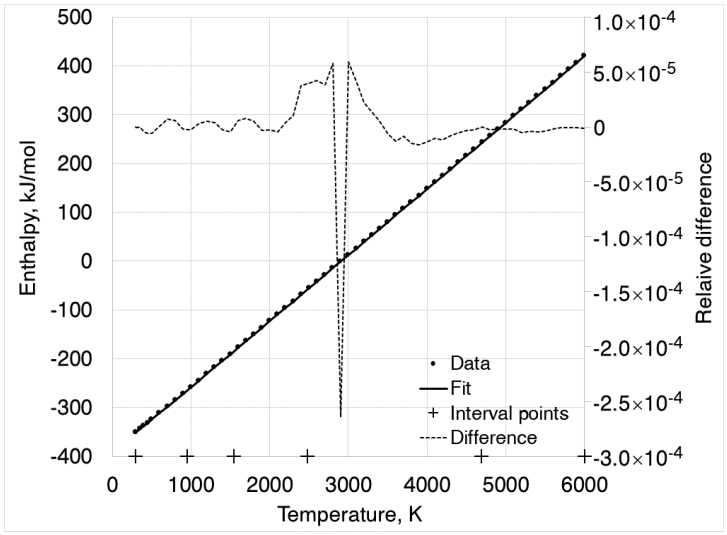


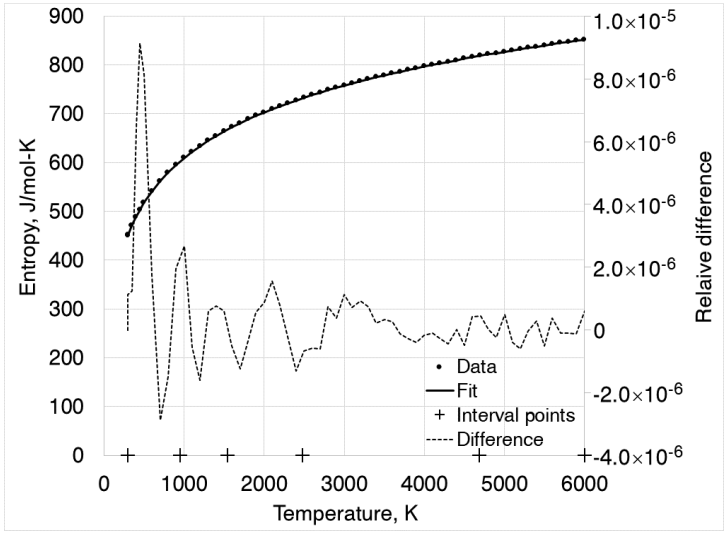


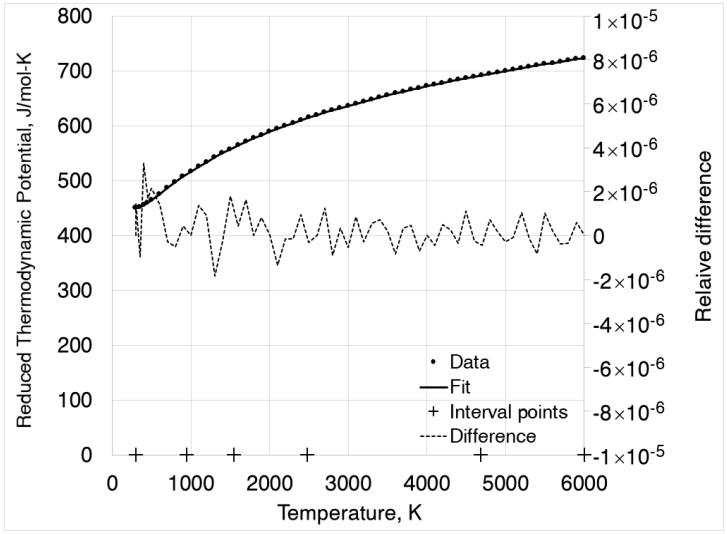
Cobalt Chloride ($(CoCl_2)_2$) – Cl4Co2(g)

MW = 259.674 g/mol – $\Delta_f H^o$(298.15 K) = -350.619 kJ/mol – S^o(298.15 K) = 450.4 J/mol-K – p^o = 0.1 MPa



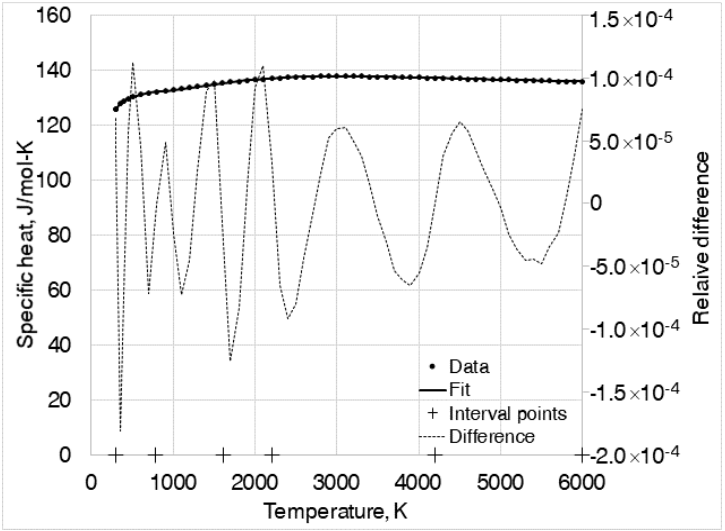


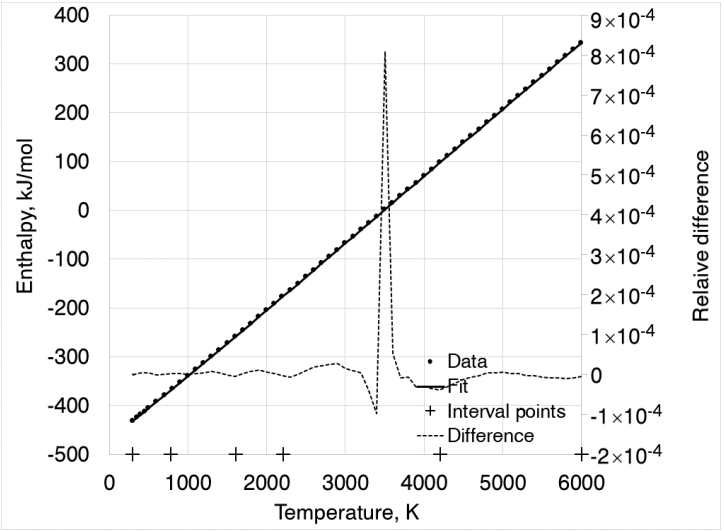


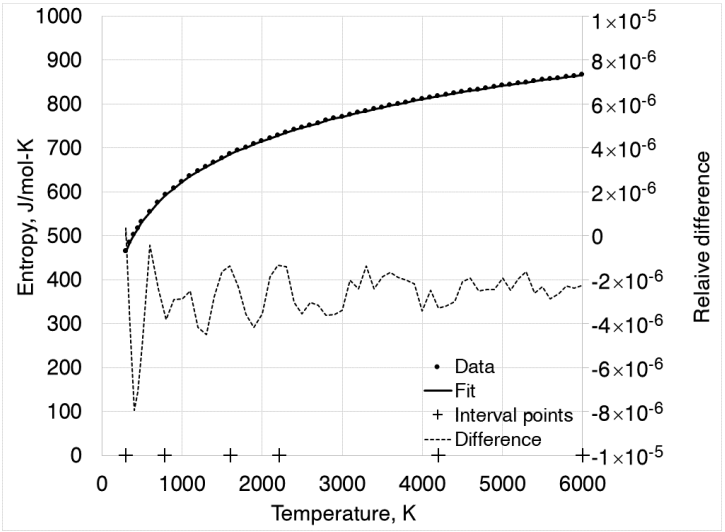


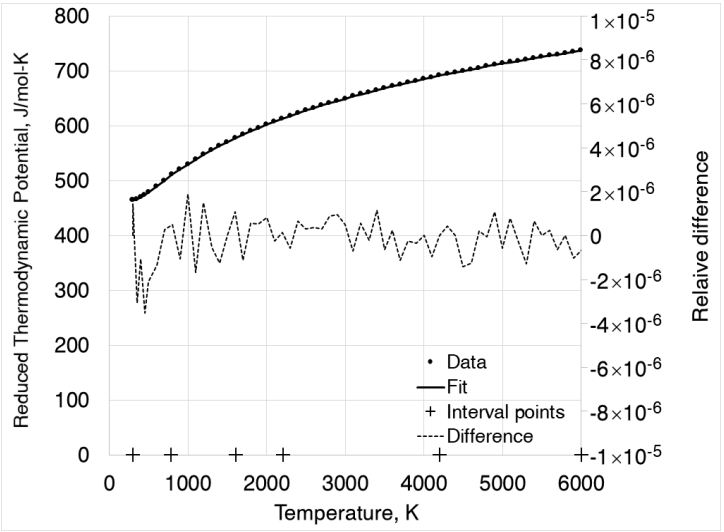
Iron Chloride ($(FeCl_2)_2$) – Cl4Fe2(g)

MW = 253.506 g/mol – $\Delta_f H^o$(298.15 K) = -431.37 kJ/mol – S^o(298.15 K) = 464.503 J/mol-K – p^o = 0.1 MPa



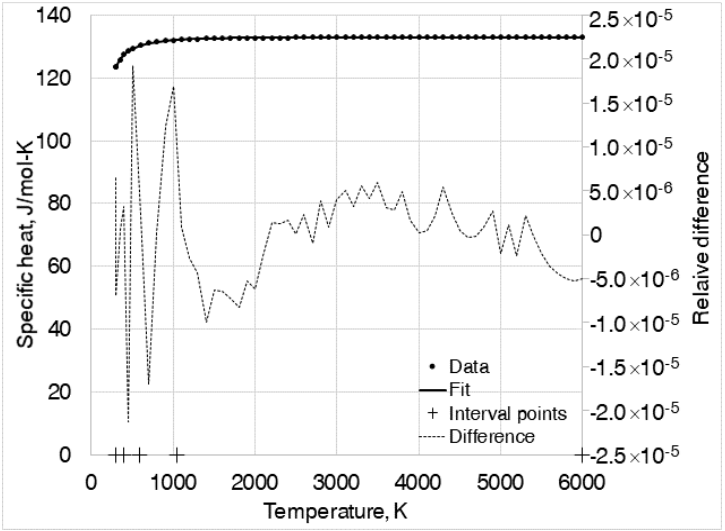


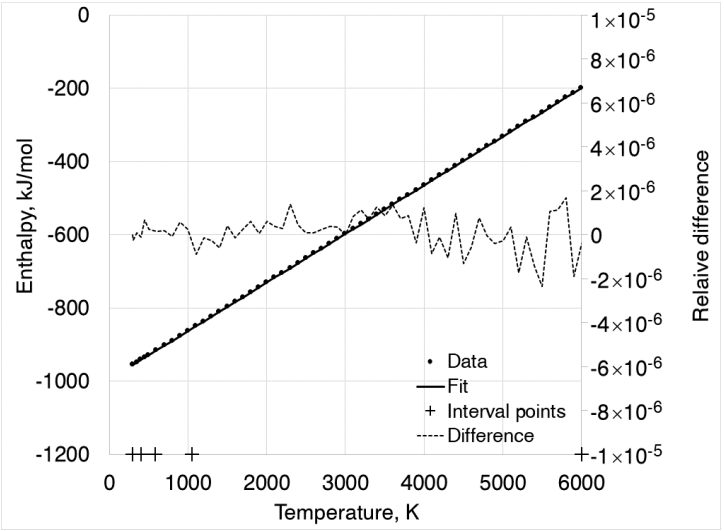


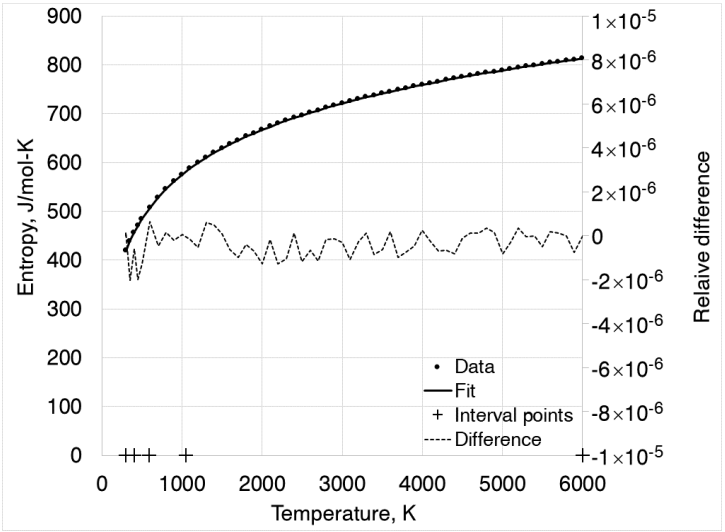


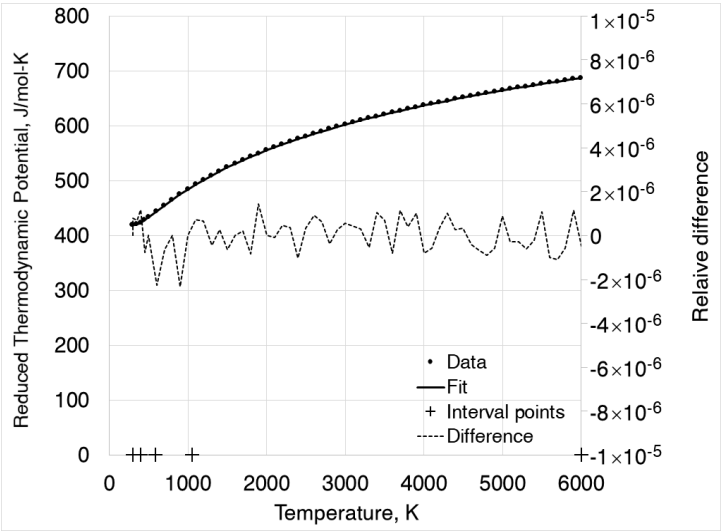
Magnesium Chloride ($(MgCl_2)_2$) – Cl4Mg2(g)

MW = 190.422 g/mol – $\Delta_f H^o$(298.15 K) = -954.37 kJ/mol – S^o(298.15 K) = 418.931 J/mol-K – p^o = 0.1 MPa



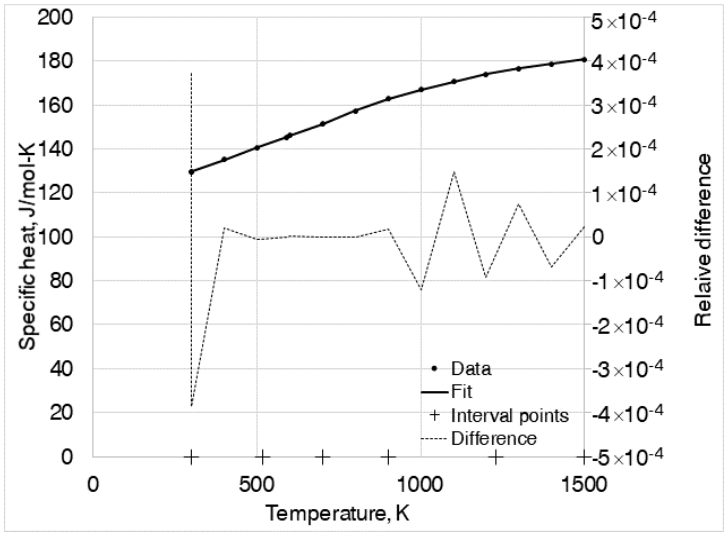


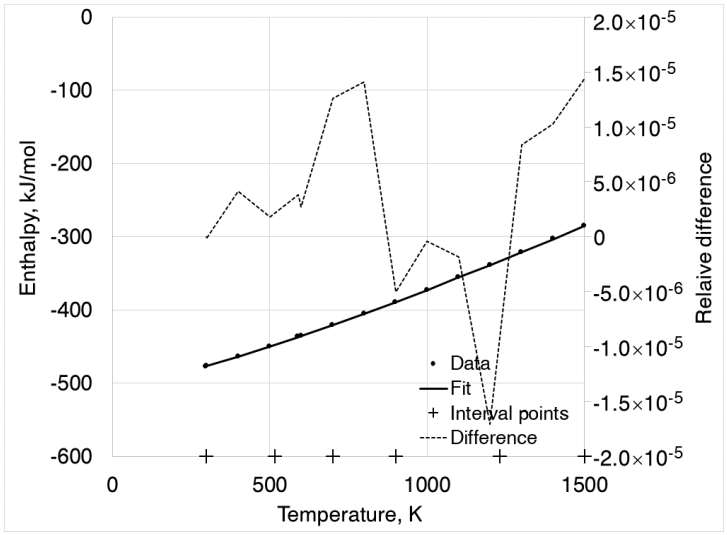


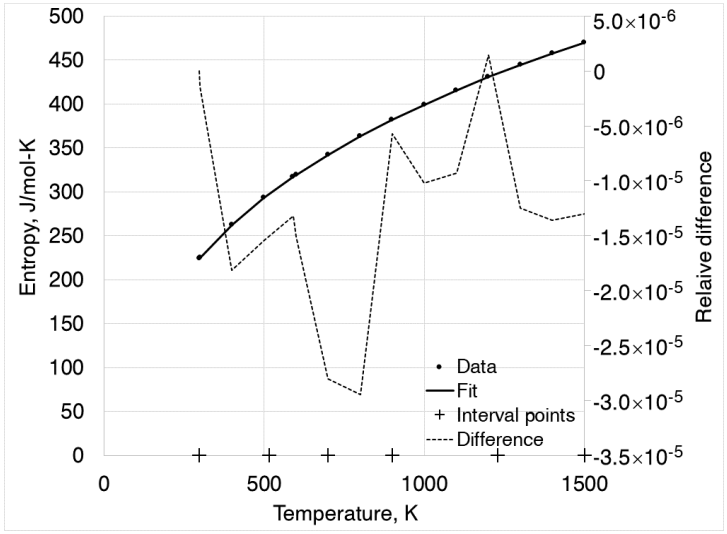


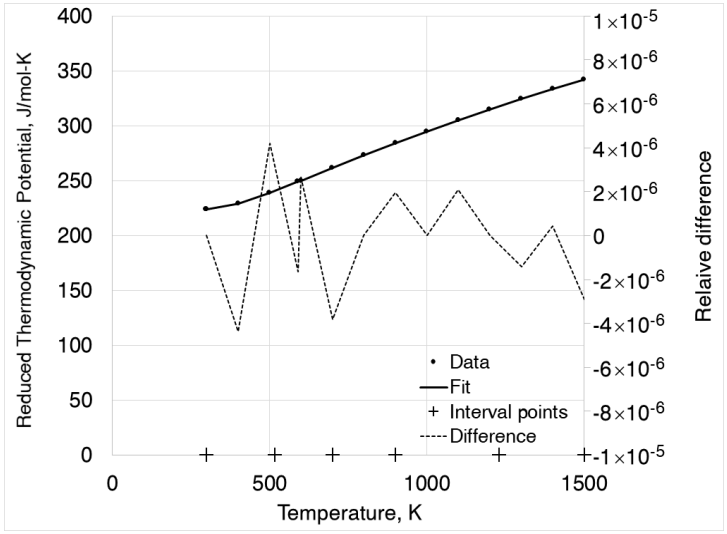
Molybdenum Chloride ($MoCl_4$) – Cl4Mo1(cr)

MW = 237.75 g/mol – $\Delta_f H^o$(298.15 K) = -476.976 kJ/mol – S^o(298.15 K) = 223.844 J/mol-K – p^o = 0.1 MPa



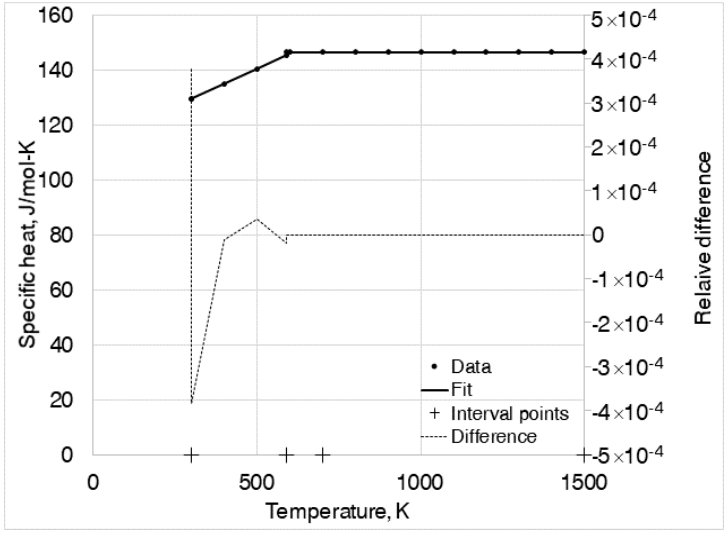


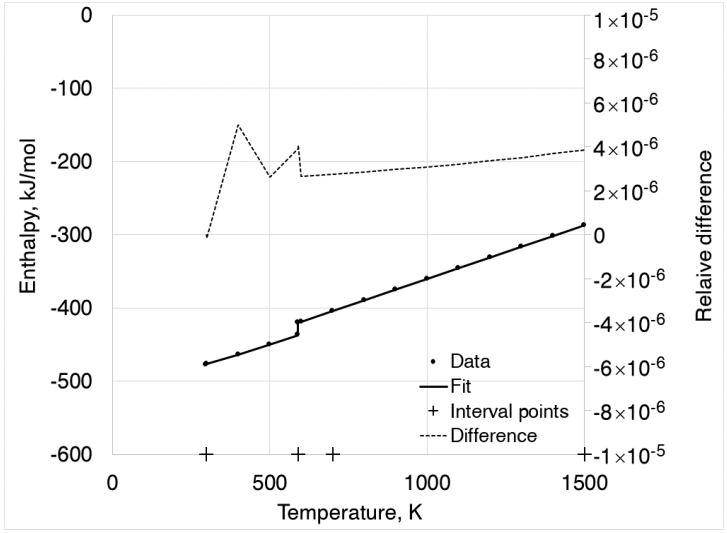


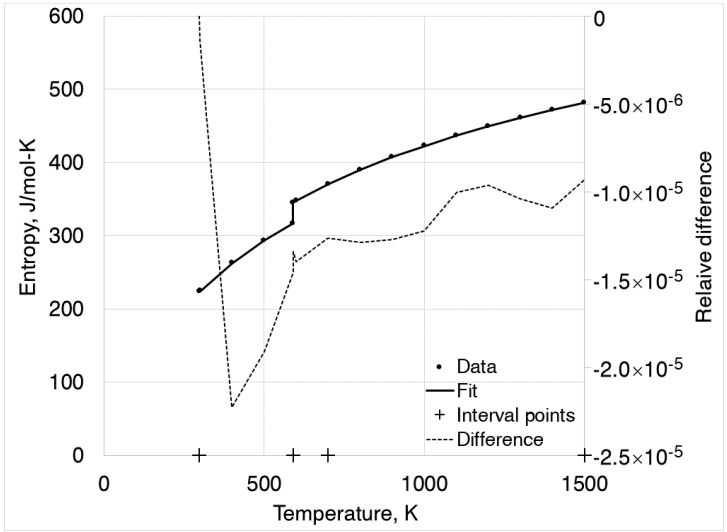


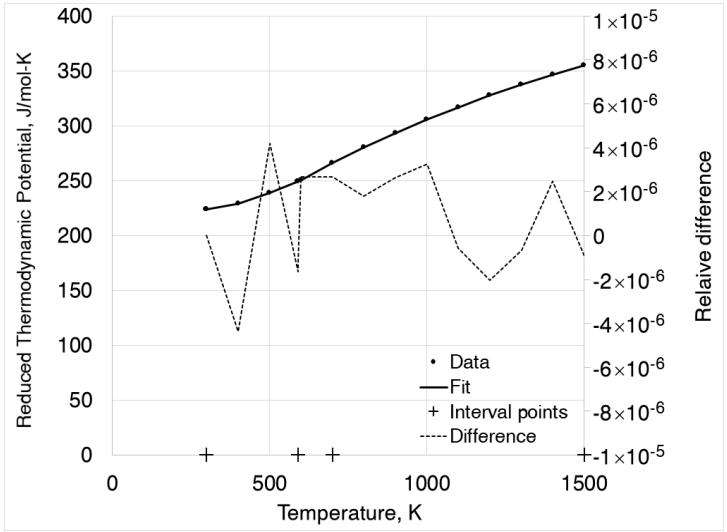
Molybdenum Chloride ($MoCl_4$) – Cl4Mo1(cr,l)

MW = 237.75 g/mol – $\Delta_f H^o$(298.15 K) = -476.976 kJ/mol – S^o(298.15 K) = 223.844 J/mol-K – p^o = 0.1 MPa



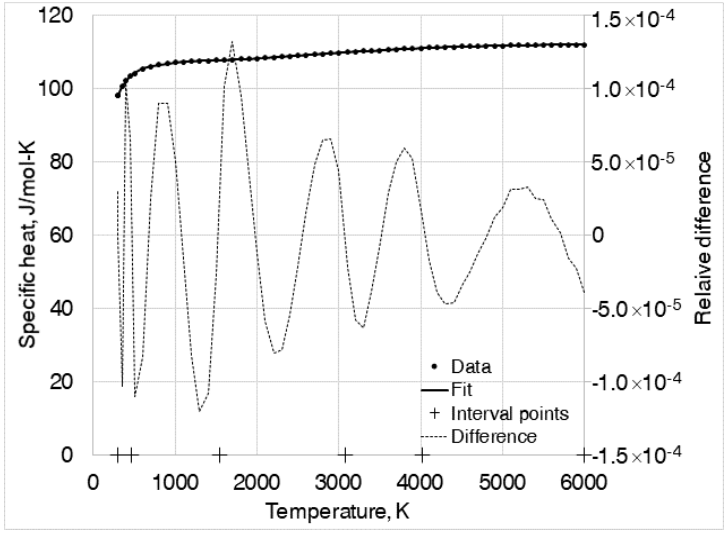


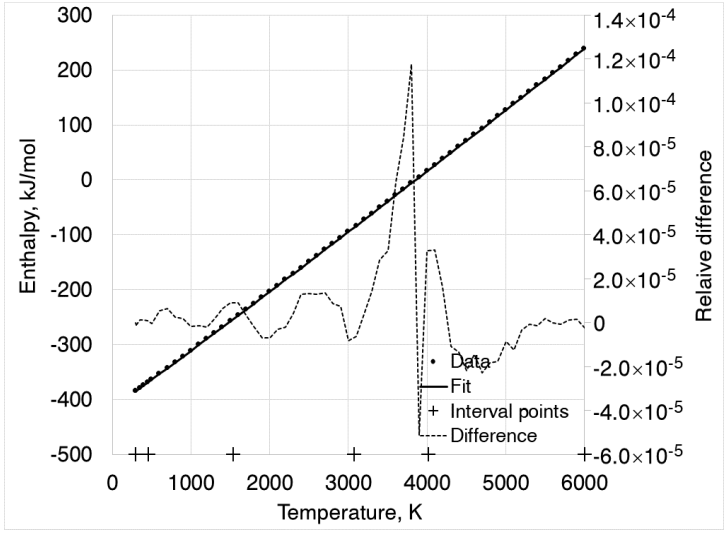


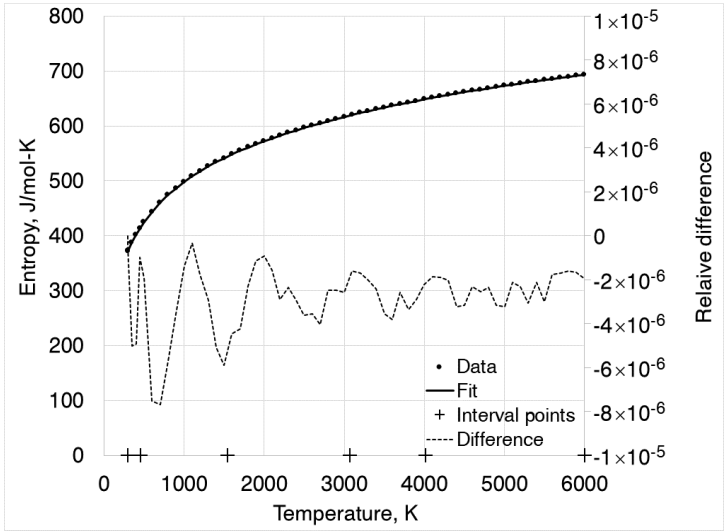


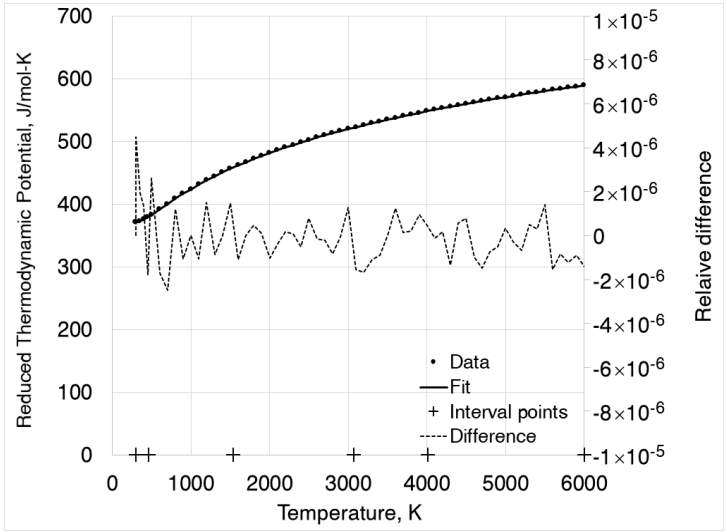
Molybdenum Chloride ($MoCl_4$) – Cl4Mo1(g)

MW = 237.75 g/mol – $\Delta_f H^o$(298.15 K) = -384.928 kJ/mol – S^o(298.15 K) = 371.941 J/mol-K – p^o = 0.1 MPa



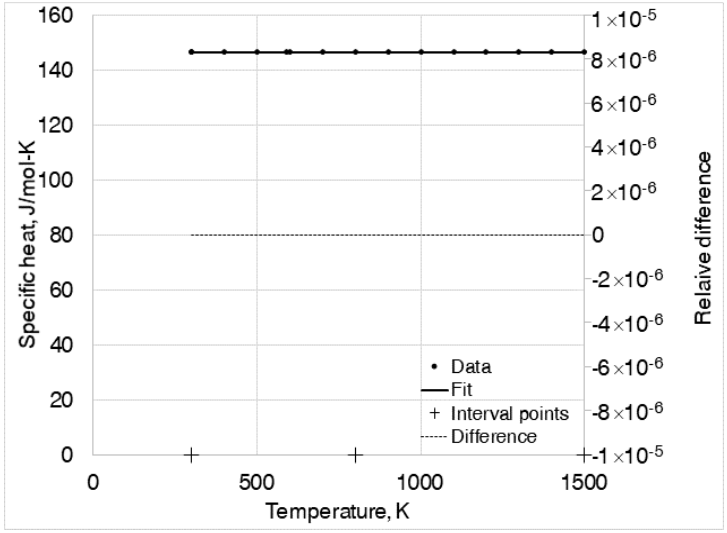


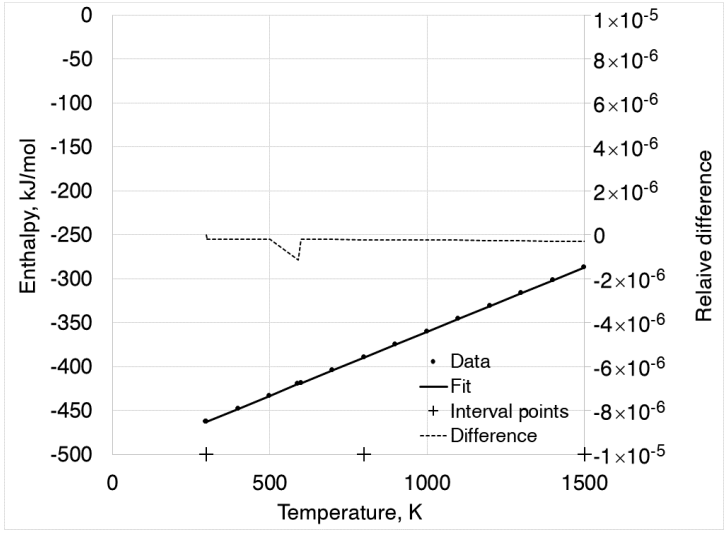


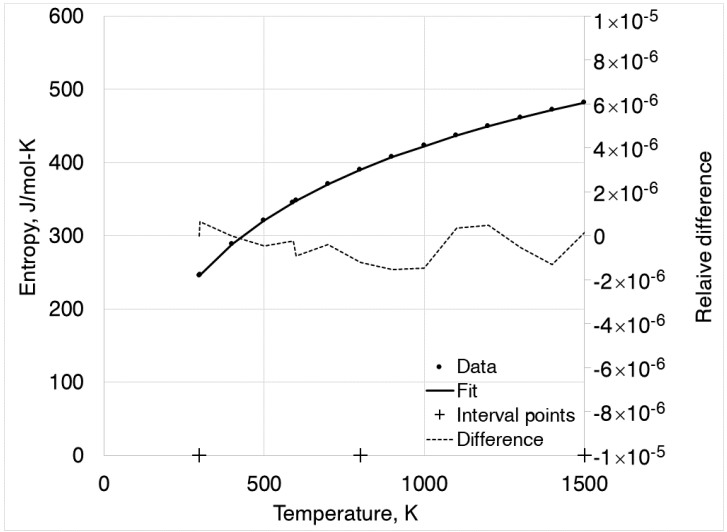


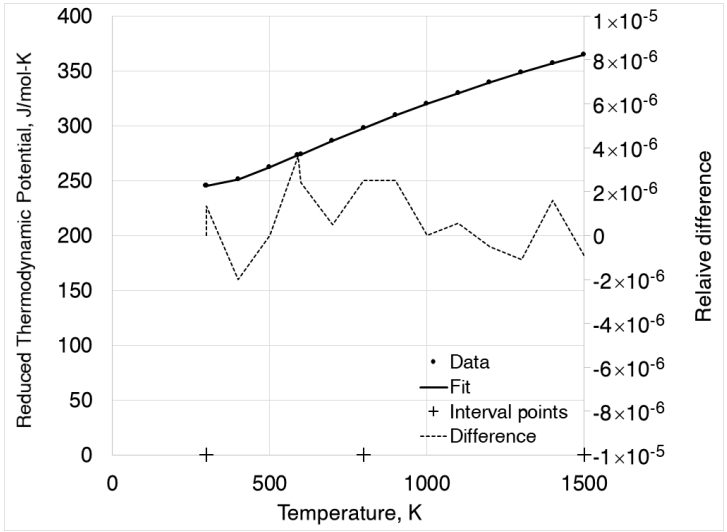
Molybdenum Chloride ($MoCl_4$) – Cl4Mo1(l)

MW = 237.75 g/mol – $\Delta_f H^o$(298.15 K) = -462.837 kJ/mol – S^o(298.15 K) = 245.525 J/mol-K – p^o = 0.1 MPa



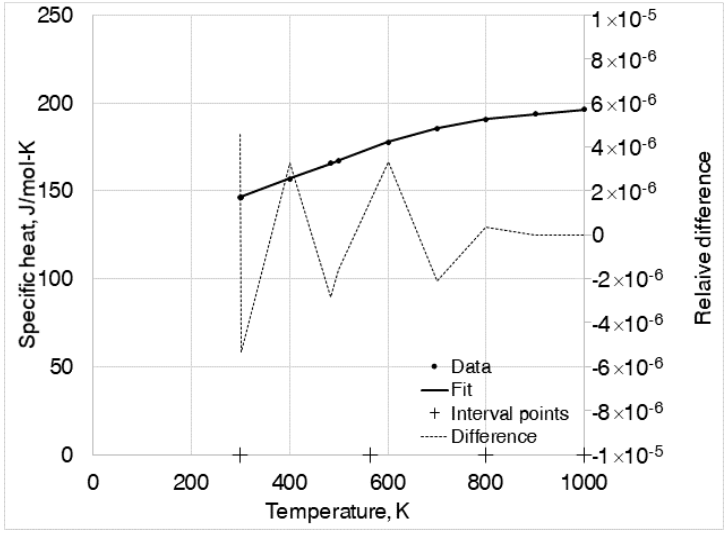


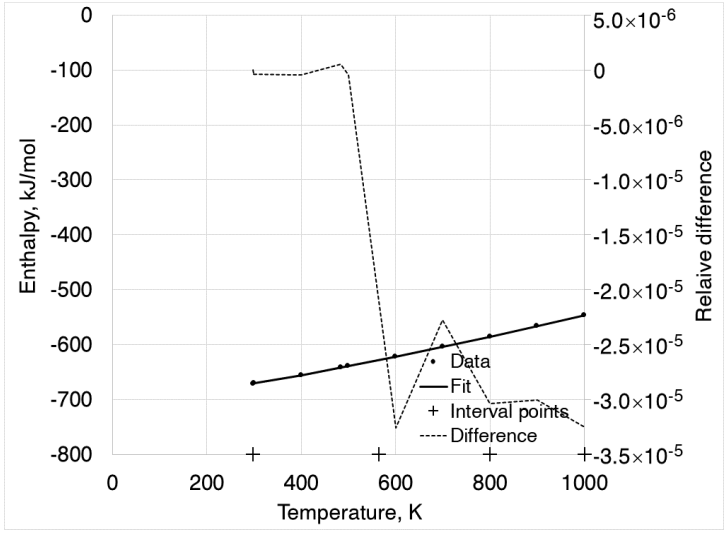


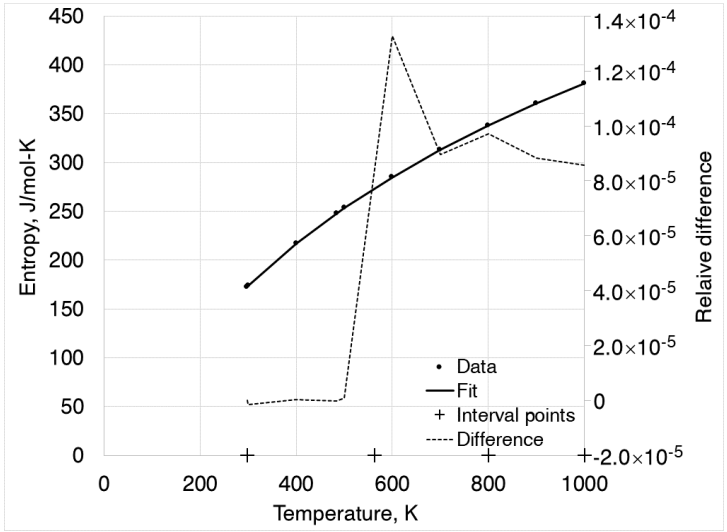


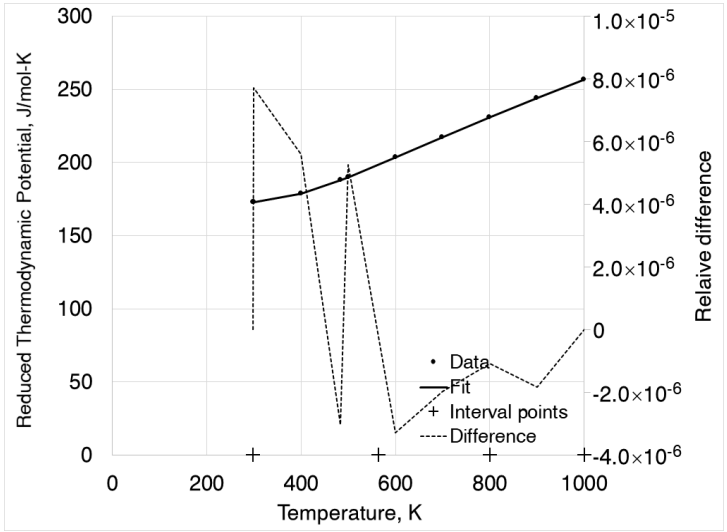
Tungsten Chloride Oxide ($OWCl_4$) – Cl4O1W1(cr)

MW = 341.66 g/mol – $\Delta_f H^o$(298.15 K) = -671.114 kJ/mol – S^o(298.15 K) = 172.799 J/mol-K – p^o = 0.1 MPa



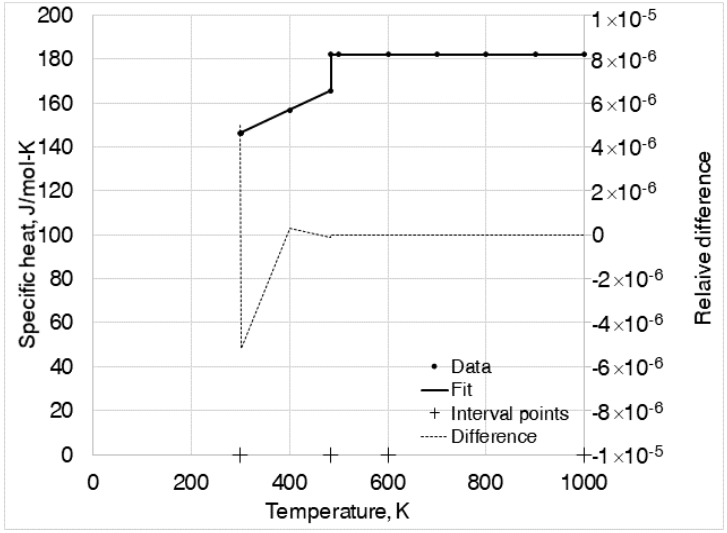


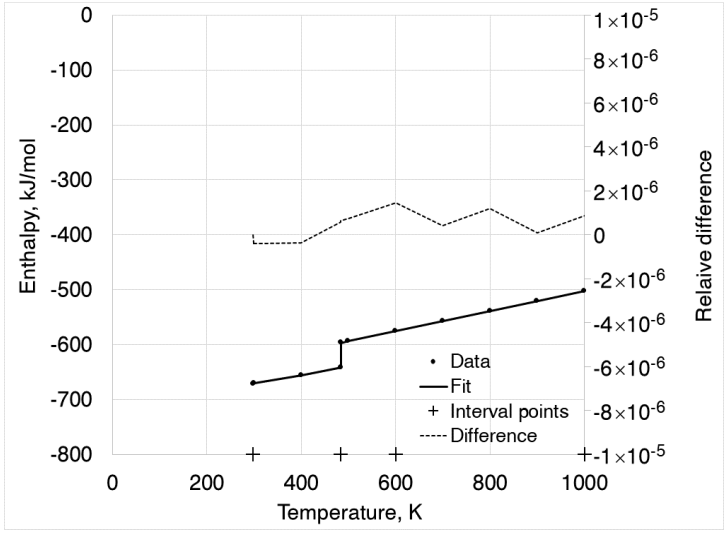


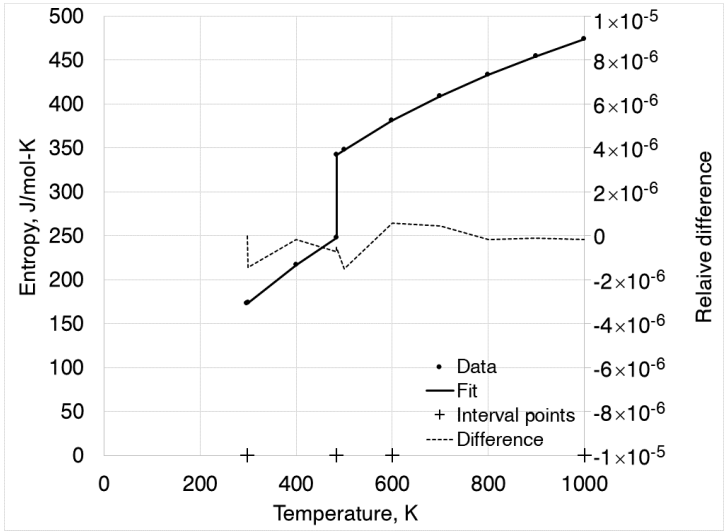


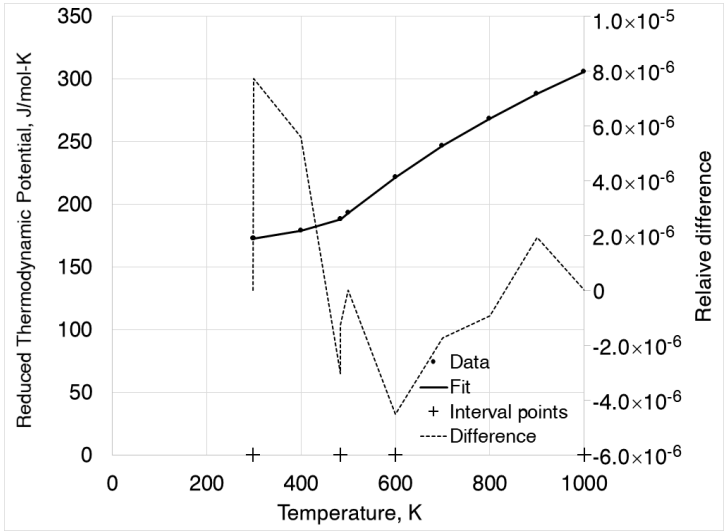
Tungsten Chloride Oxide ($OWCl_4$) – Cl4O1W1(cr,l)

MW = 341.66 g/mol – $\Delta_f H^o$(298.15 K) = -671.114 kJ/mol – S^o(298.15 K) = 172.799 J/mol-K – p^o = 0.1 MPa



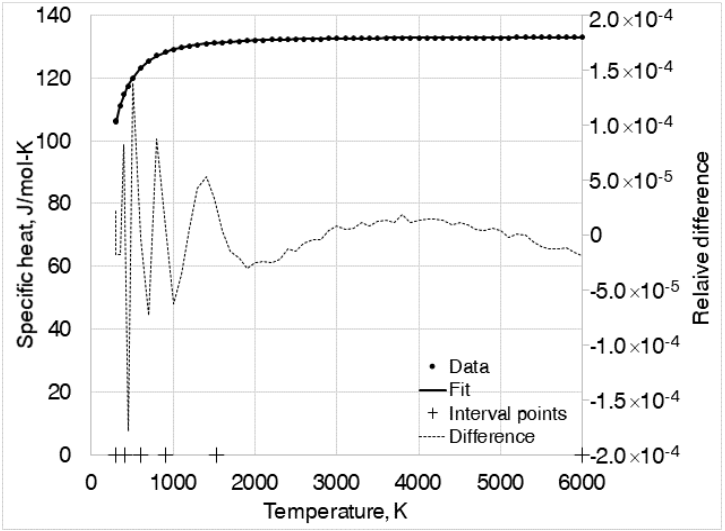


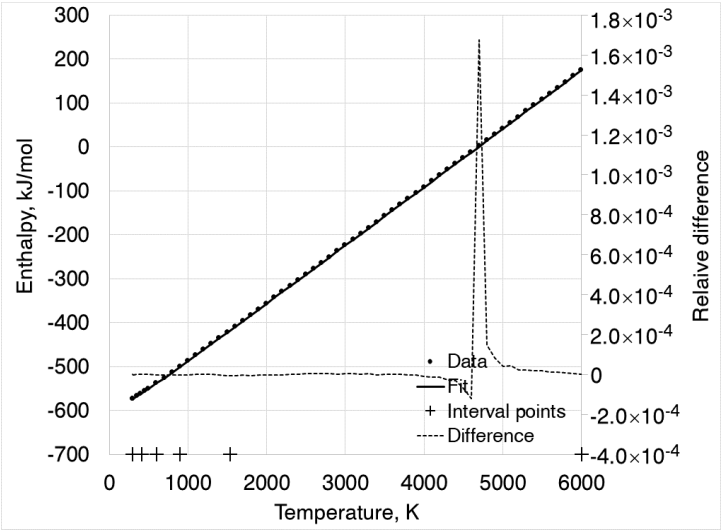


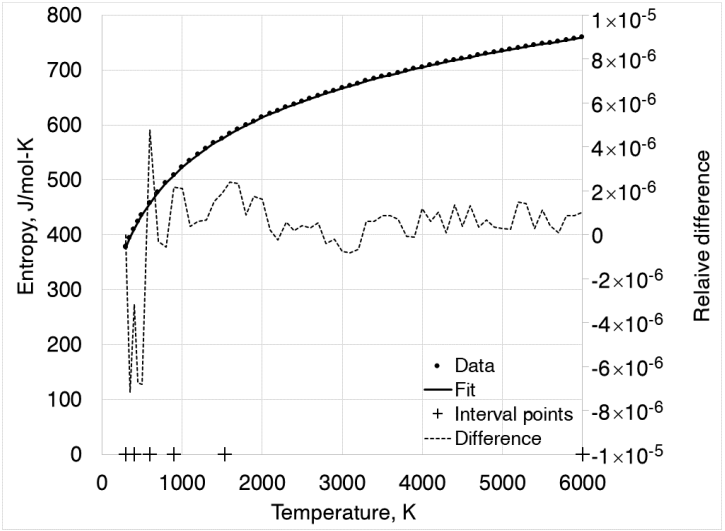


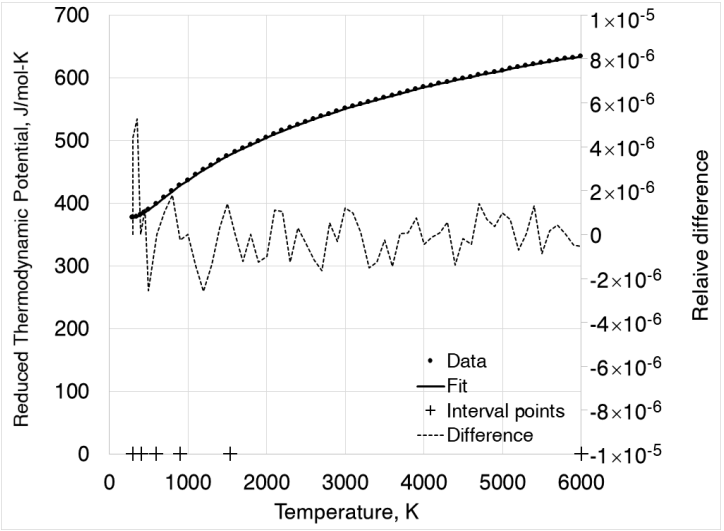
Tungsten Chloride Oxide ($OWCl_4$) – Cl4O1W1(g)

MW = 341.66 g/mol – $\Delta_f H^o$(298.15 K) = -573.493 kJ/mol – S^o(298.15 K) = 377.079 J/mol-K – p^o = 0.1 MPa



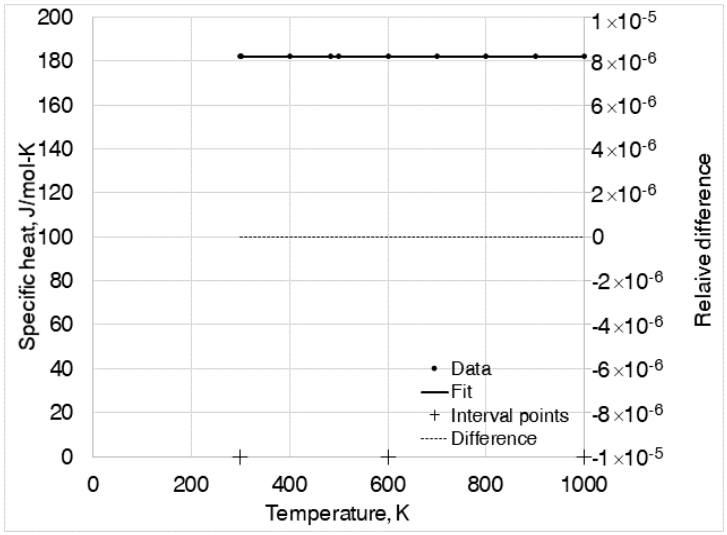


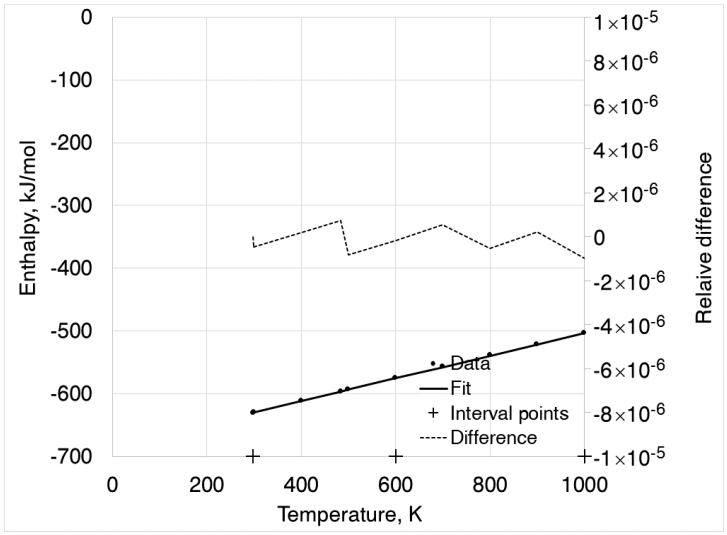


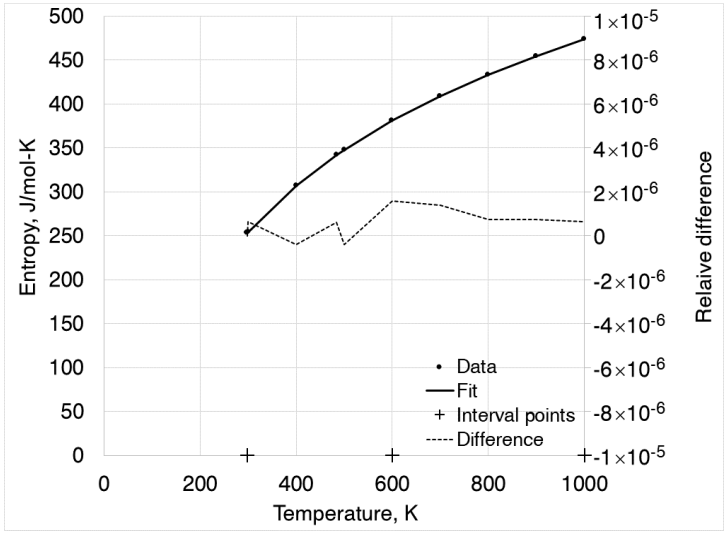


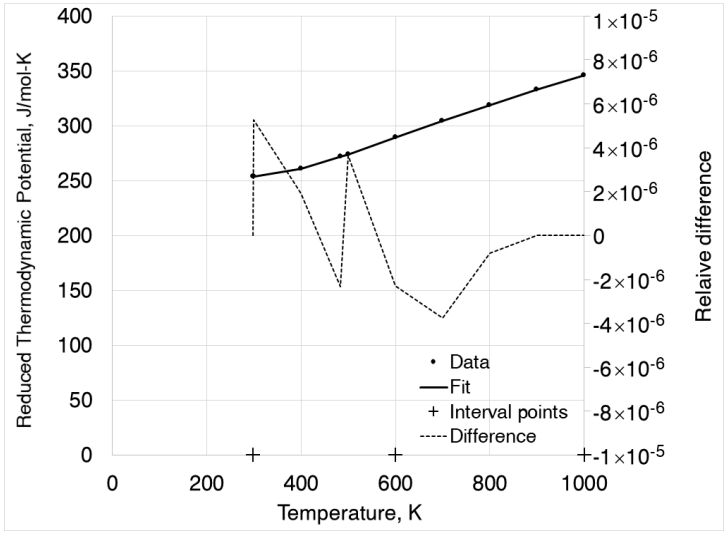
Tungsten Chloride Oxide ($OWCl_4$) – Cl4O1W1(l)

MW = 341.66 g/mol – $\Delta_f H^o$(298.15 K) = -630.573 kJ/mol – S^o(298.15 K) = 253.565 J/mol-K – p^o = 0.1 MPa



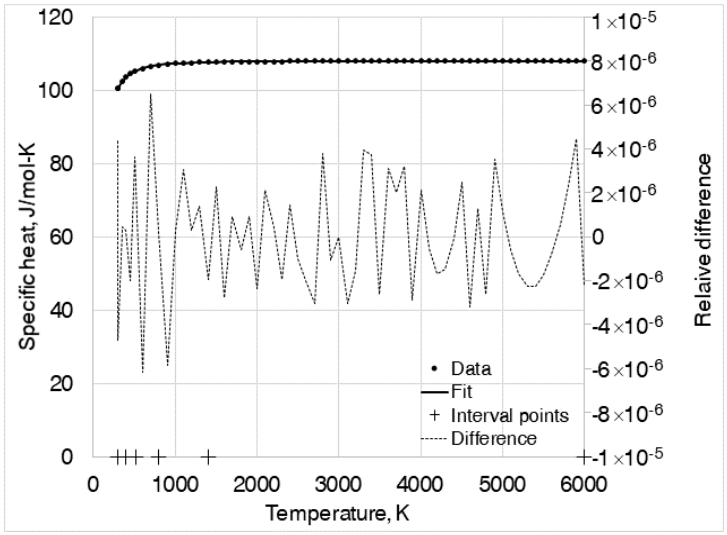


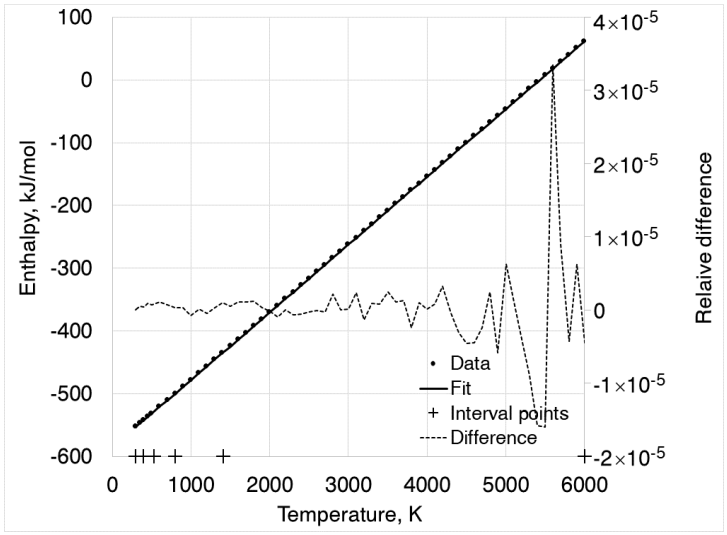


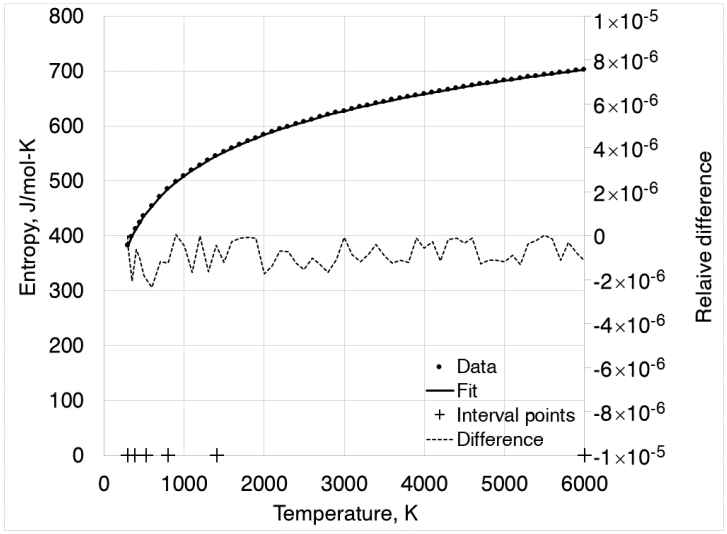


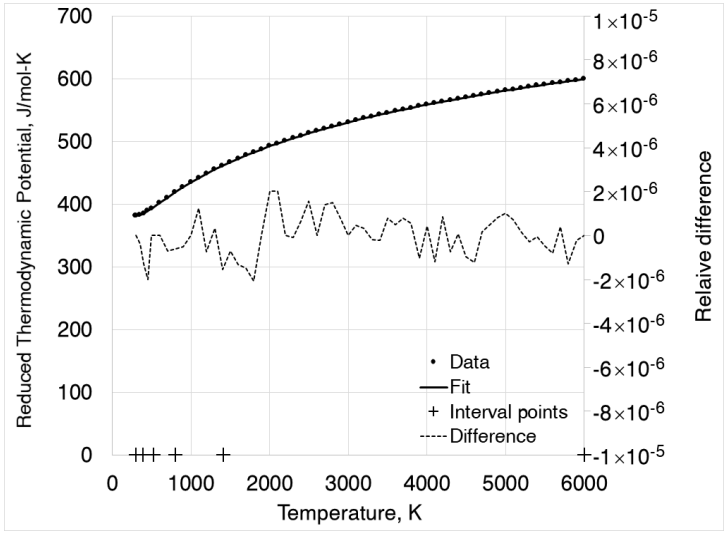
Lead Chloride ($PbCl_4$) – Cl4Pb1(g)

MW = 349 g/mol – $\Delta_f H^o$(298.15 K) = -552.403 kJ/mol – S^o(298.15 K) = 381.659 J/mol-K – p^o = 0.1 MPa



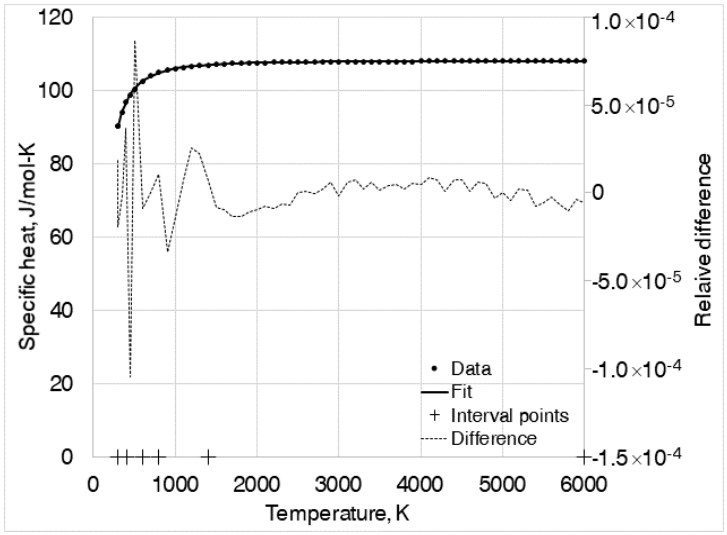


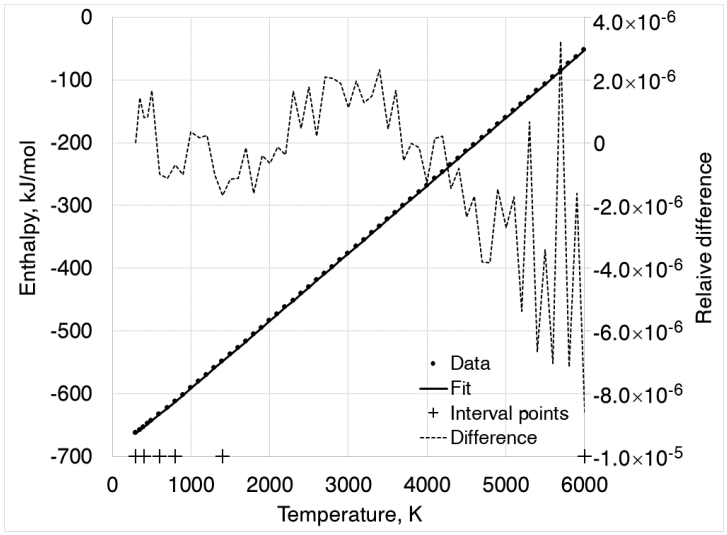


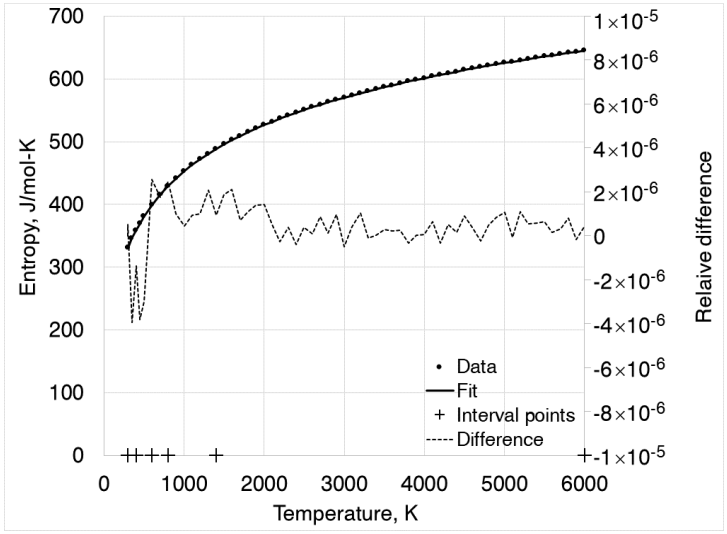


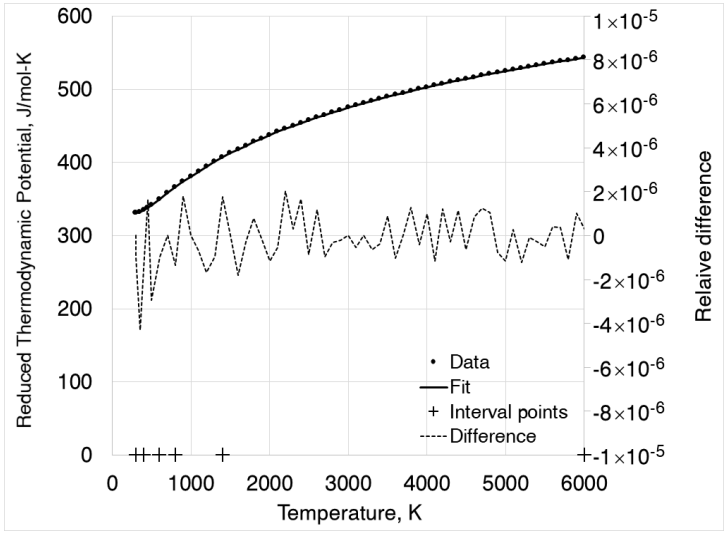
Tetrachlorosilane ($SiCl_4$) – Cl4Si1(g)

MW = 169.898 g/mol – $\Delta_f H^o$(298.15 K) = -662.746 kJ/mol – S^o(298.15 K) = 330.945 J/mol-K – p^o = 0.1 MPa



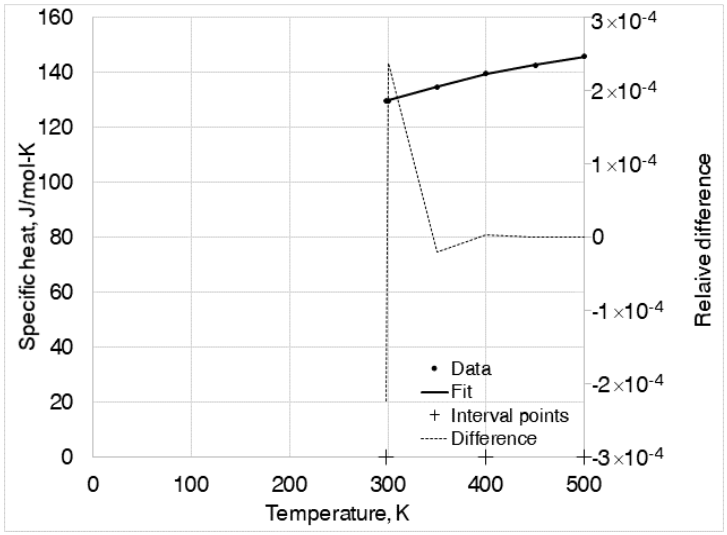


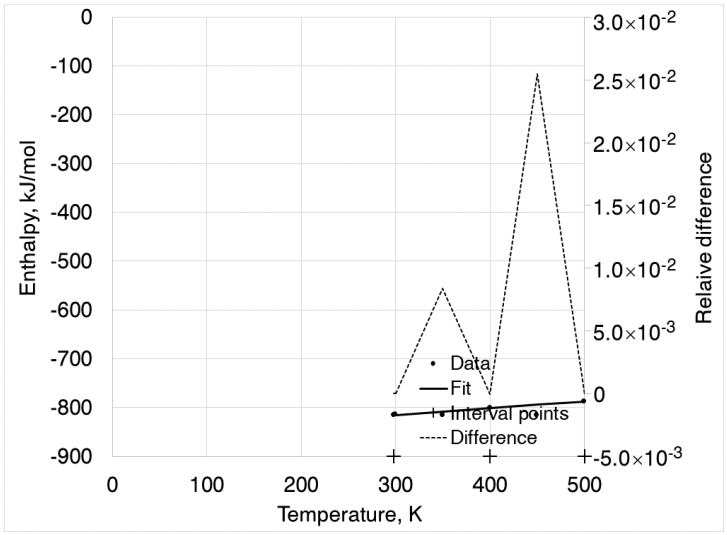


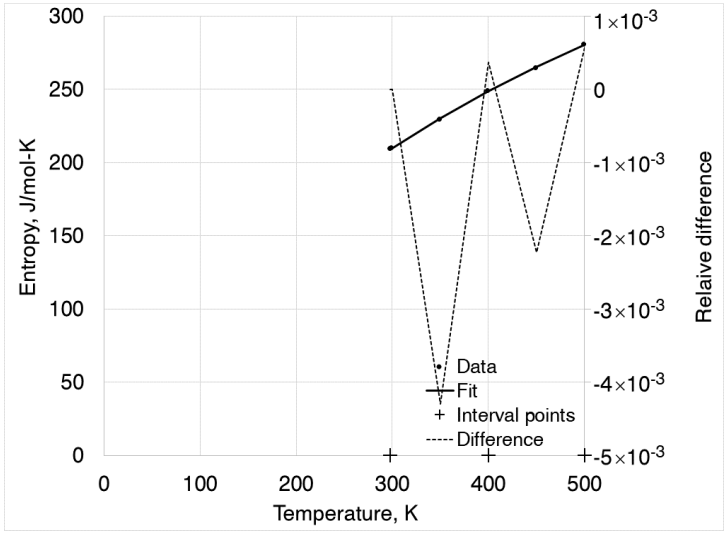


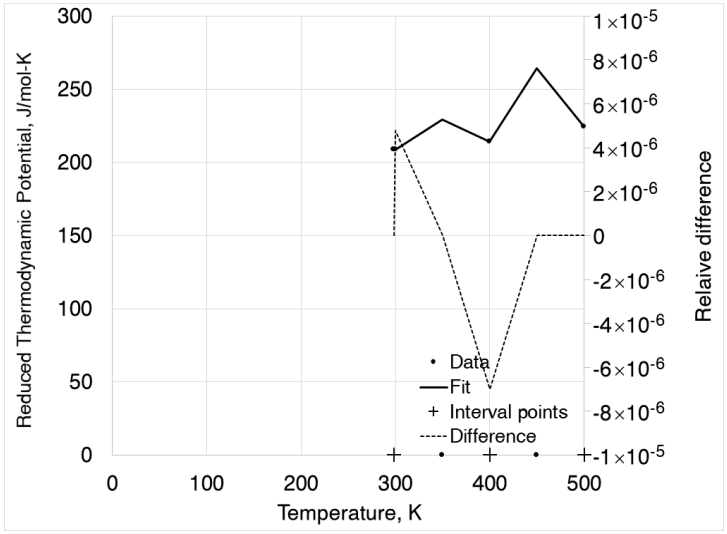
Titanium Chloride ($TiCl_4$) – Cl4Ti1(cr)

MW = 189.69 g/mol – $\Delta_f H^o$(298.15 K) = -815.026 kJ/mol – S^o(298.15 K) = 209.069 J/mol-K – p^o = 0.1 MPa



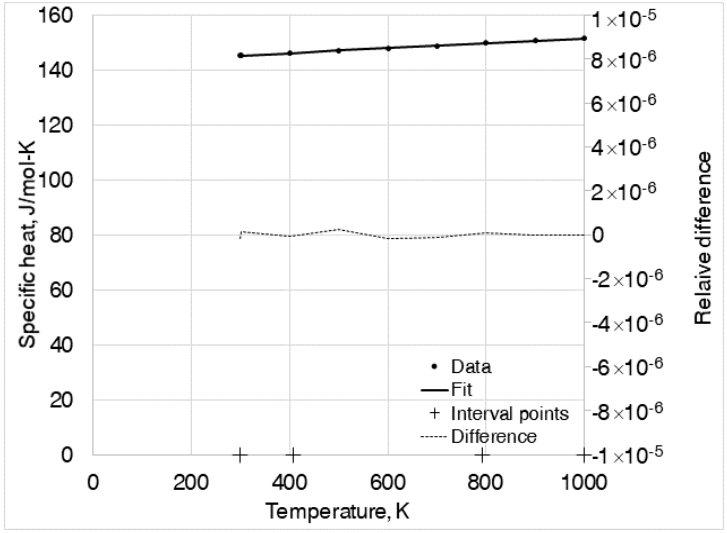


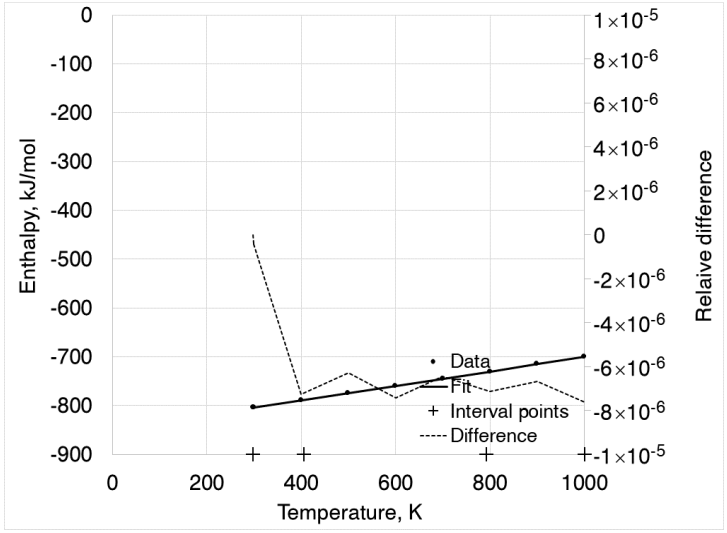


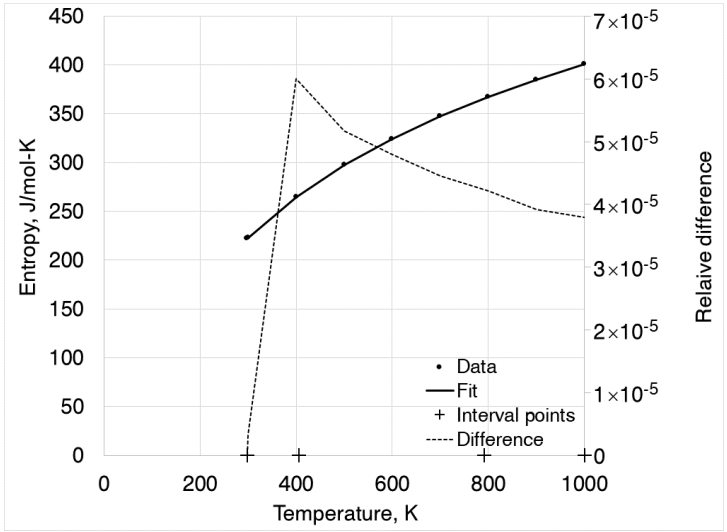


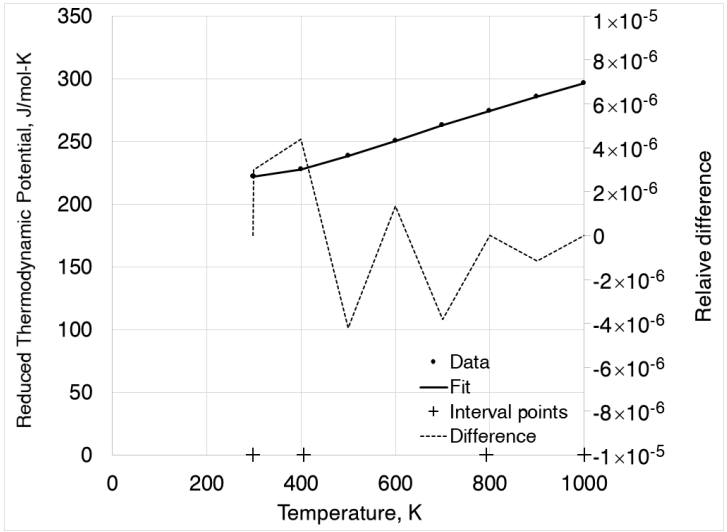
Titanium Chloride ($TiCl_4$) – Cl4Ti1(cr,l)

MW = 189.69 g/mol – $\Delta_f H^o$(298.15 K) = -804.165 kJ/mol – S^o(298.15 K) = 221.925 J/mol-K – p^o = 0.1 MPa



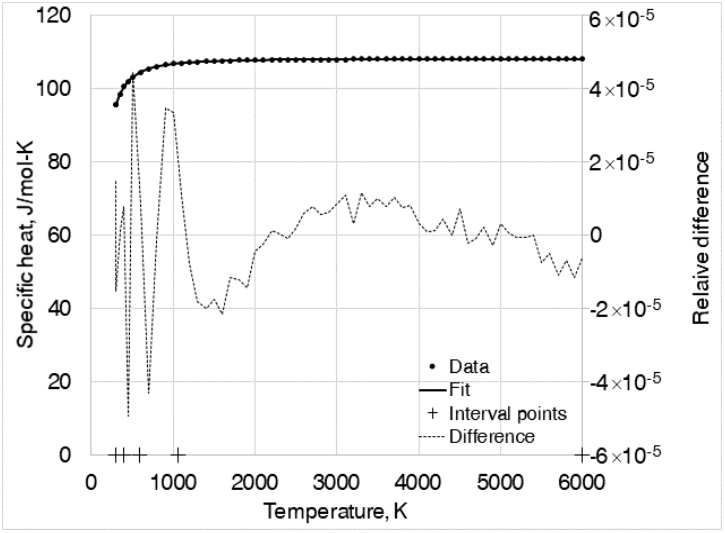


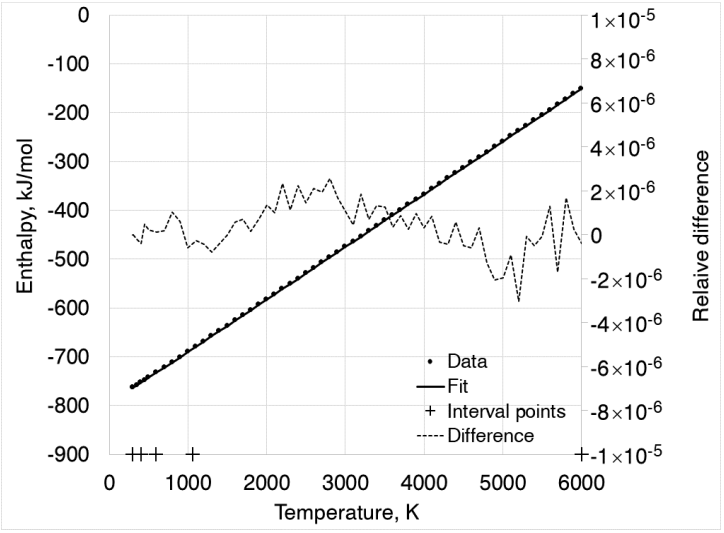


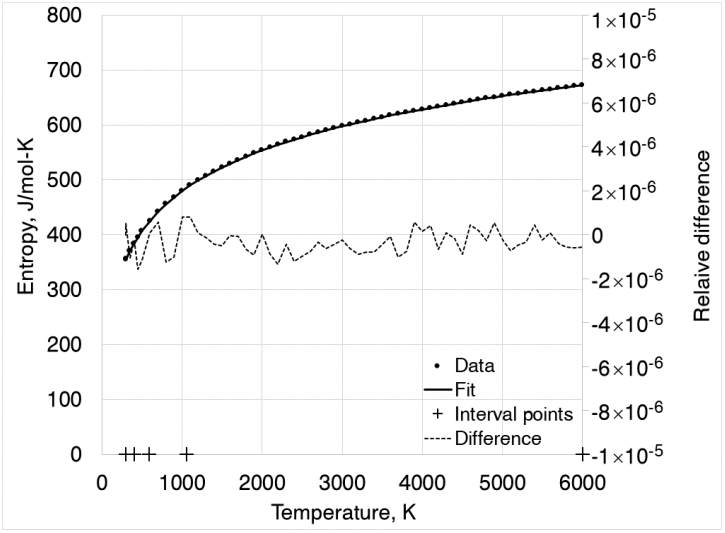


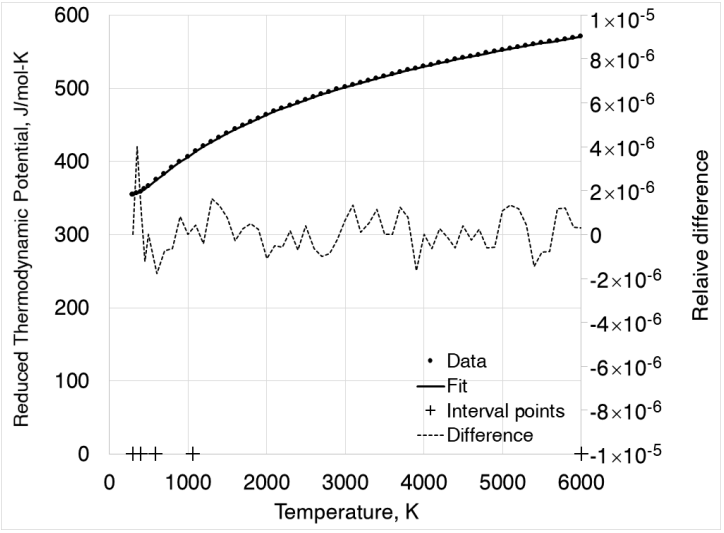
Titanium Chloride ($TiCl_4$) – Cl4Ti1(g)

MW = 189.69 g/mol – $\Delta_f H^o$(298.15 K) = -763.162 kJ/mol – S^o(298.15 K) = 354.889 J/mol-K – p^o = 0.1 MPa



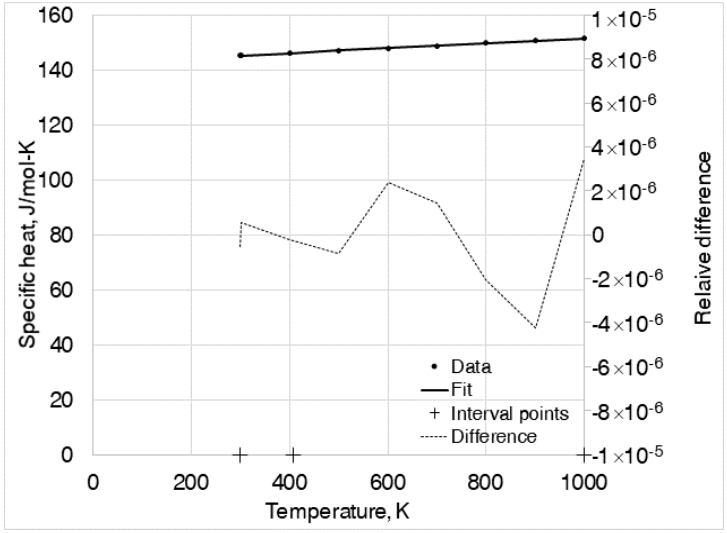


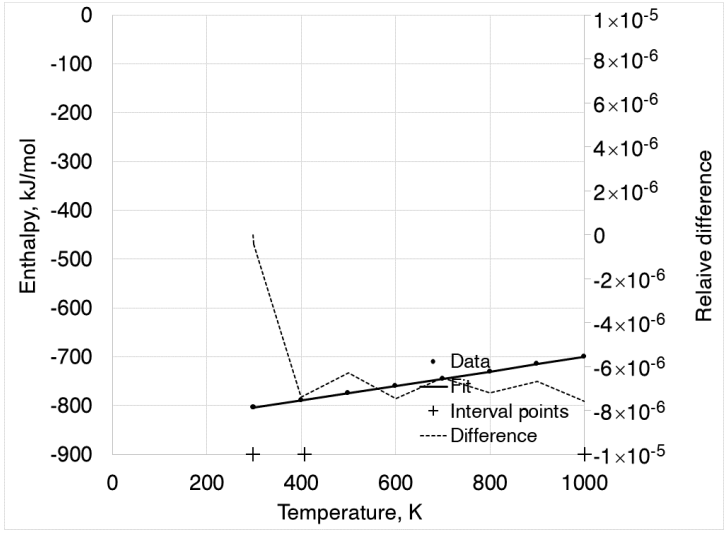


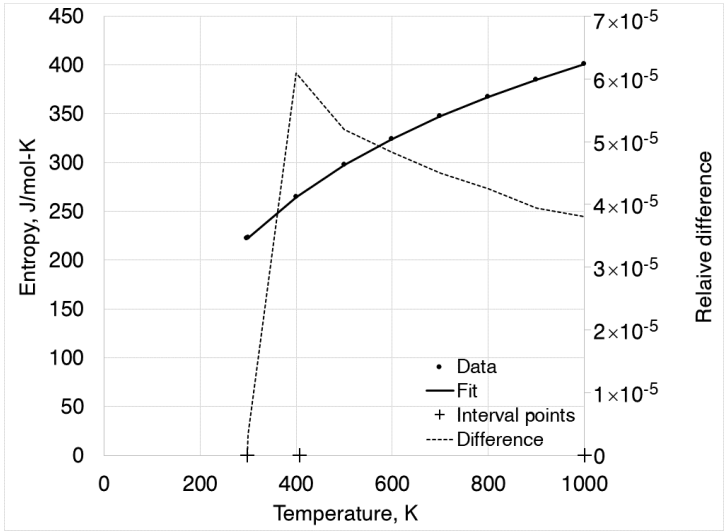


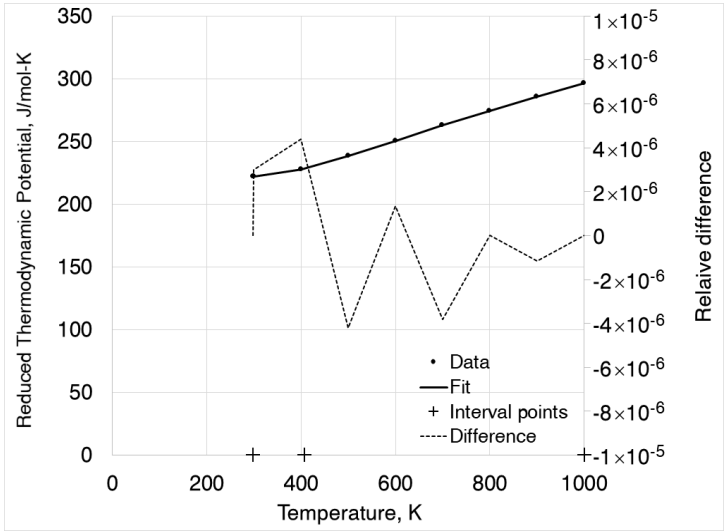
Titanium Chloride ($TiCl_4$) – Cl4Ti1(l)

MW = 189.69 g/mol – $\Delta_f H^o$(298.15 K) = -804.165 kJ/mol – S^o(298.15 K) = 221.925 J/mol-K – p^o = 0.1 MPa



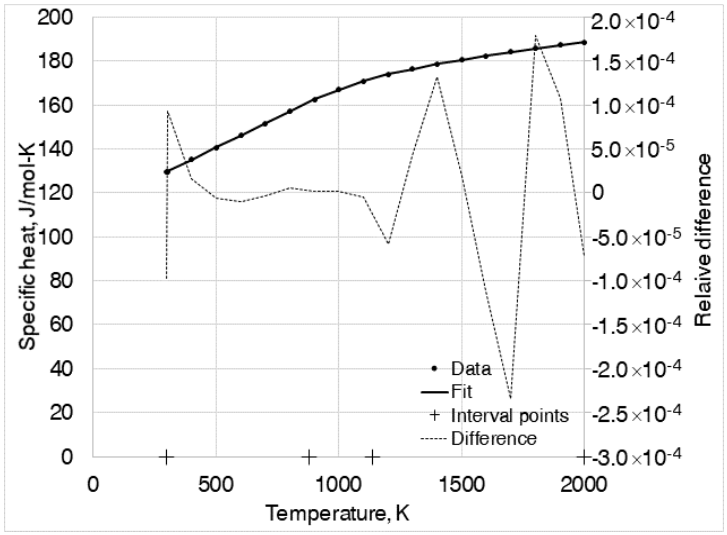


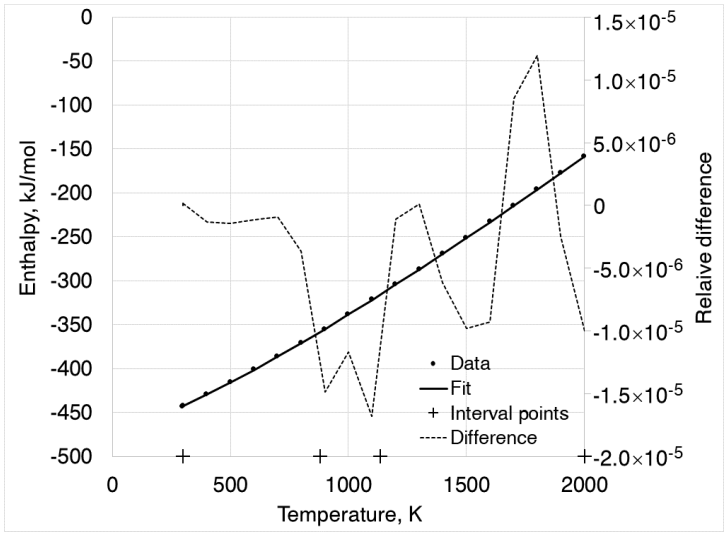


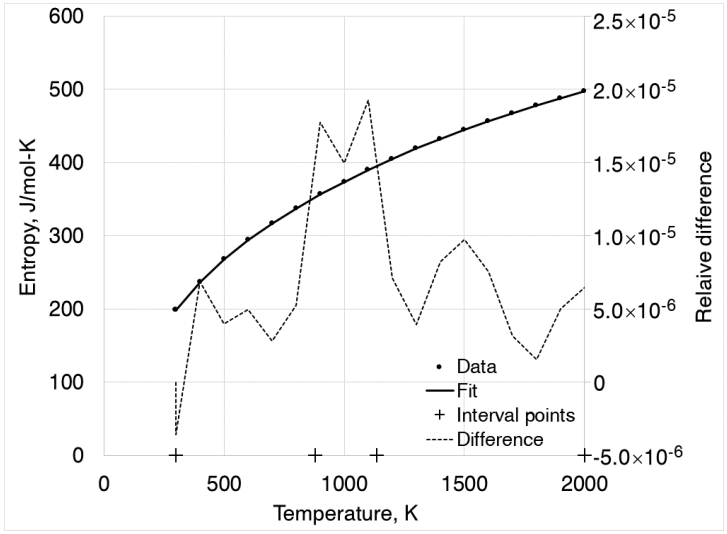


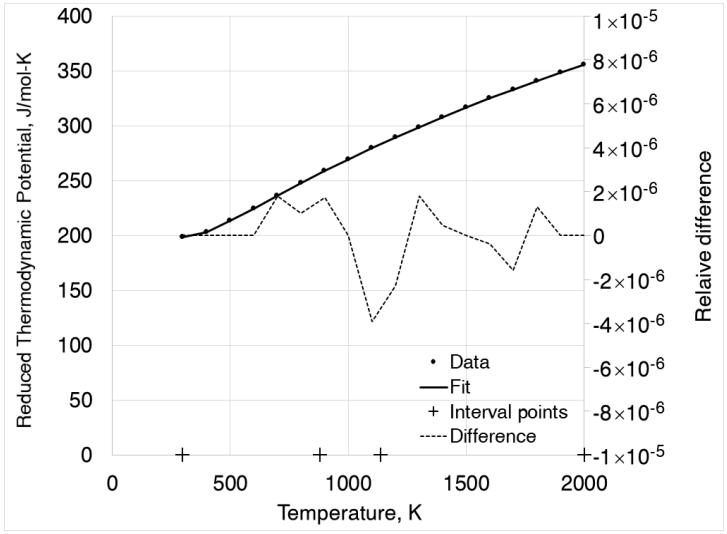
Tungsten Chloride (WCl_4) – Cl4W1(cr)

MW = 325.66 g/mol – $\Delta_f H^o$(298.15 K) = -443.086 kJ/mol – S^o(298.15 K) = 198.322 J/mol-K – p^o = 0.1 MPa



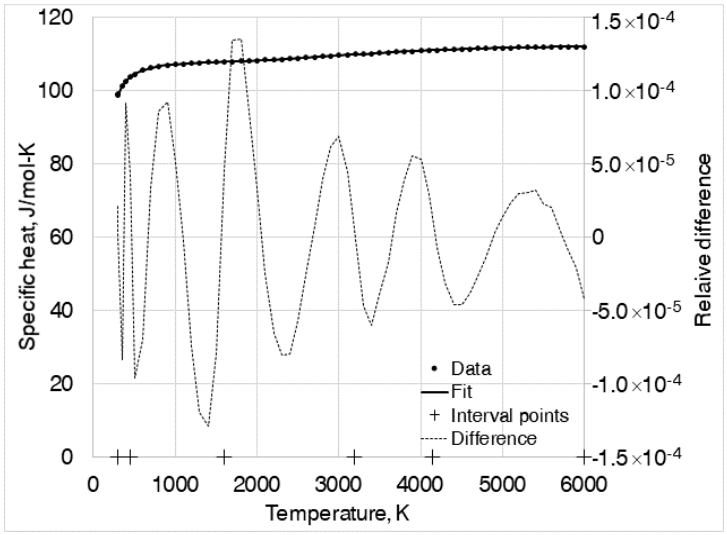


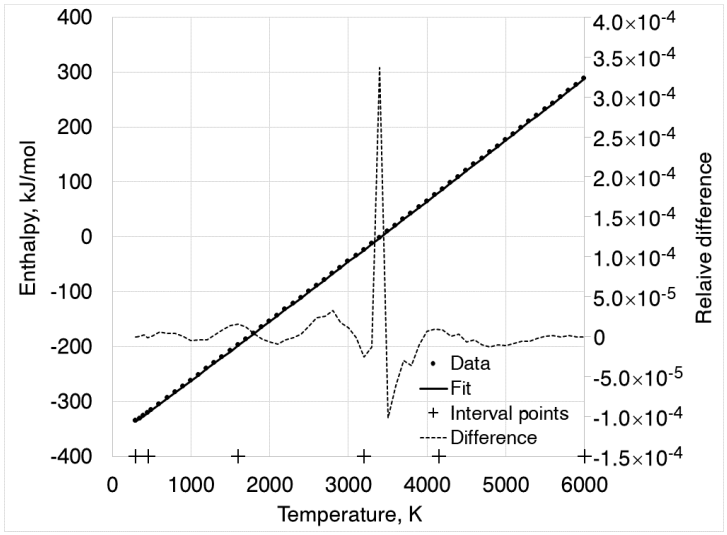


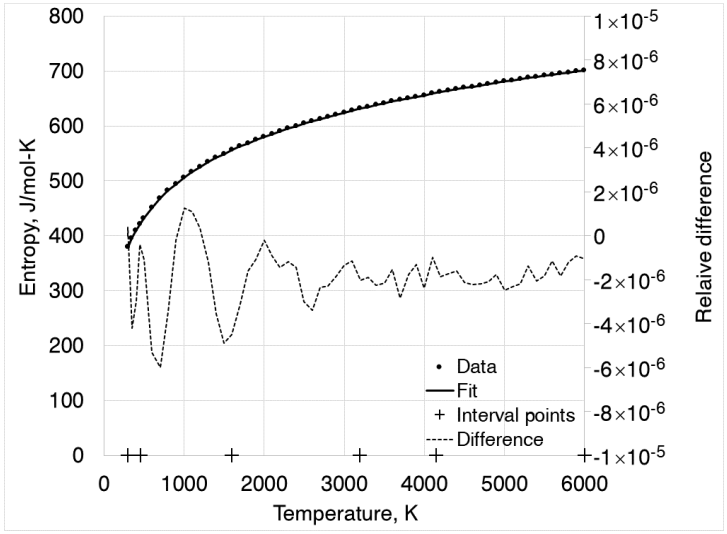


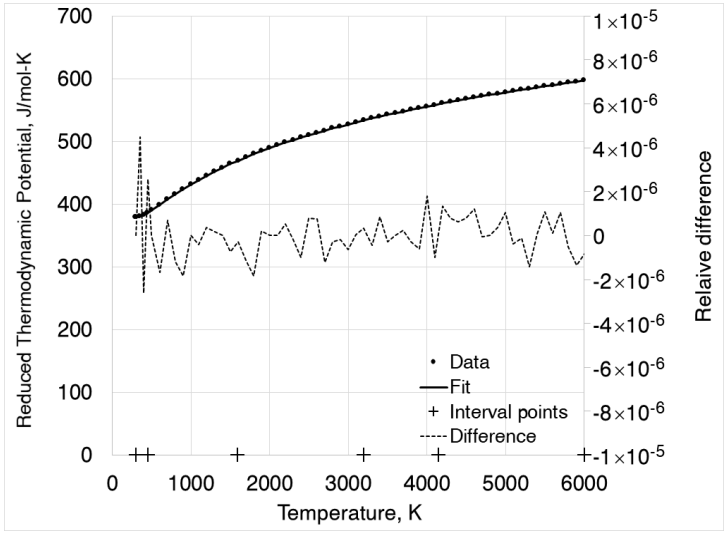
Tungsten Chloride (WCl_4) – Cl4W1(g)

MW = 325.66 g/mol – $\Delta_f H^o$(298.15 K) = -335.975 kJ/mol – S^o(298.15 K) = 379.285 J/mol-K – p^o = 0.1 MPa



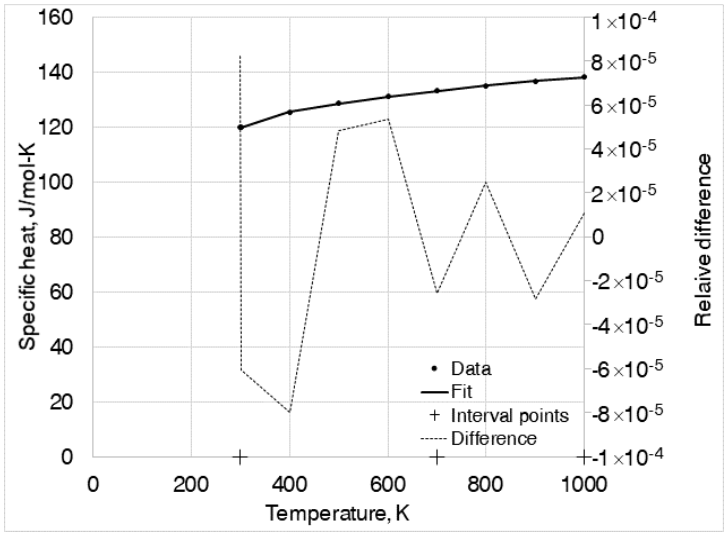


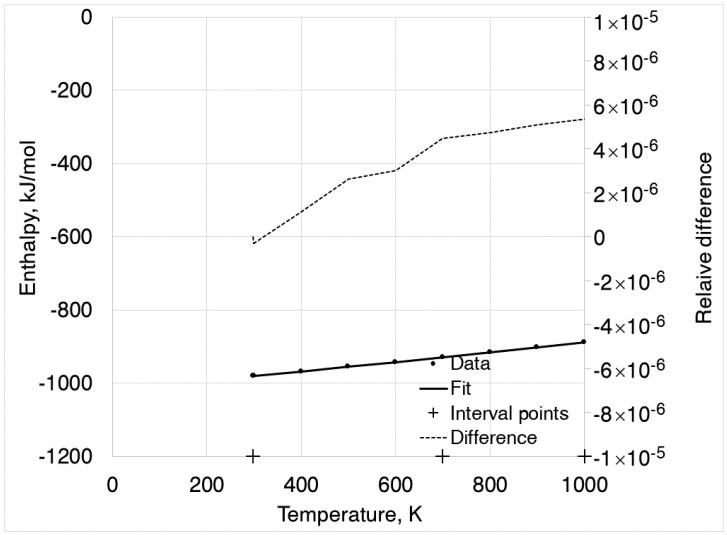


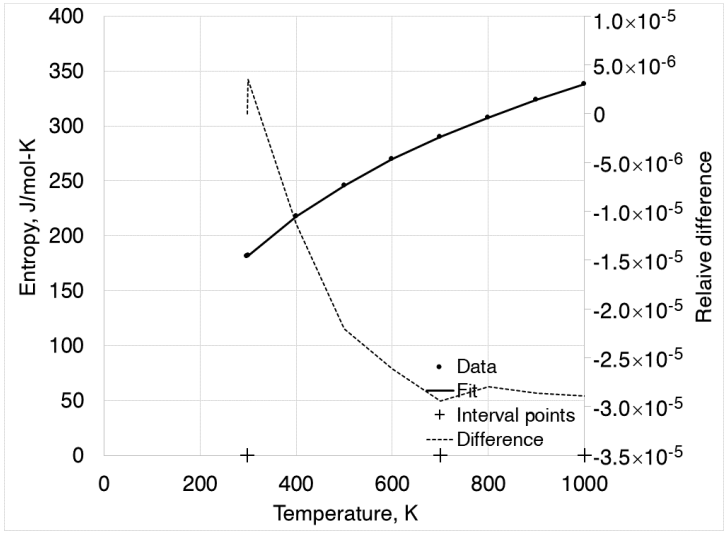


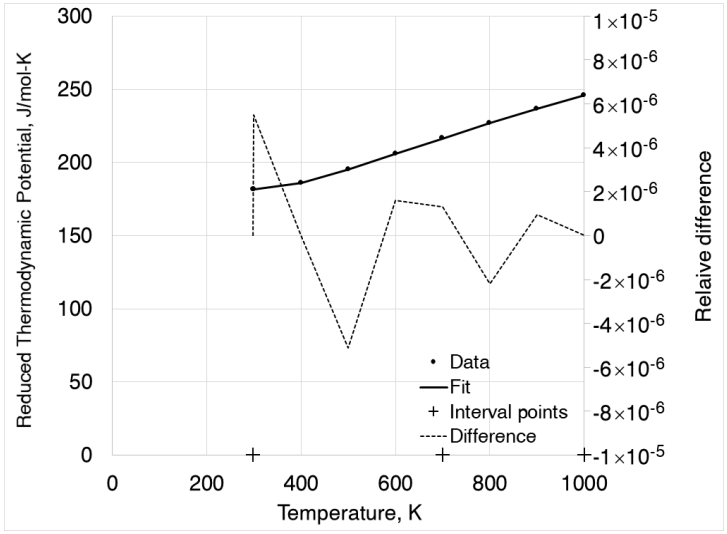
Zirconium Chloride ($ZrCl_4$) – Cl4Zr1(cr)

MW = 233.03 g/mol – $\Delta_f H^o$(298.15 K) = -980.52 kJ/mol – S^o(298.15 K) = 181.418 J/mol-K – p^o = 0.1 MPa



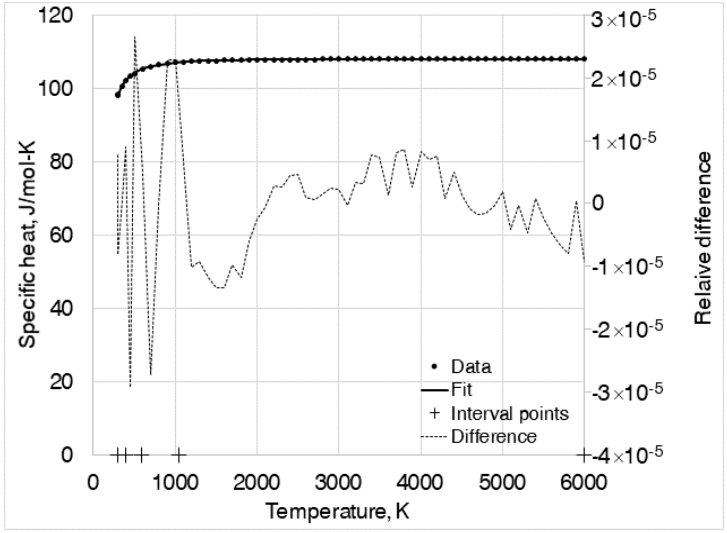


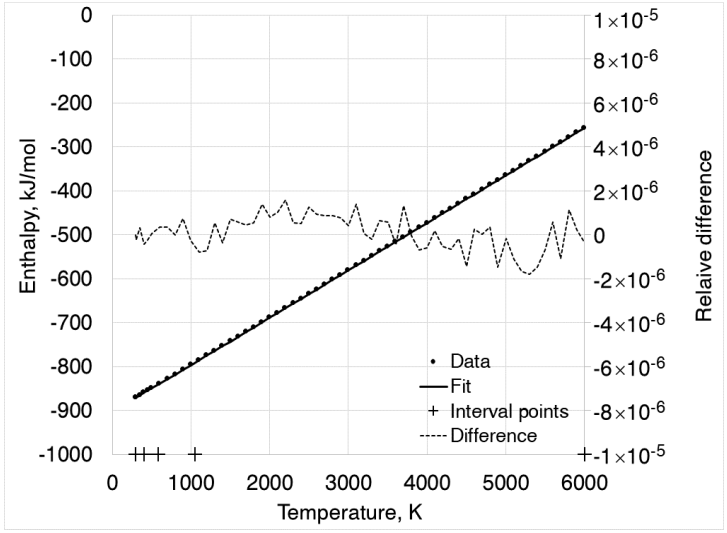


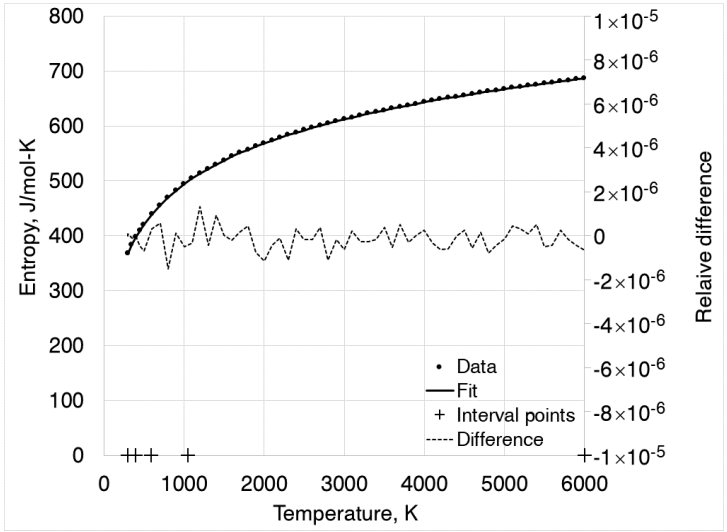


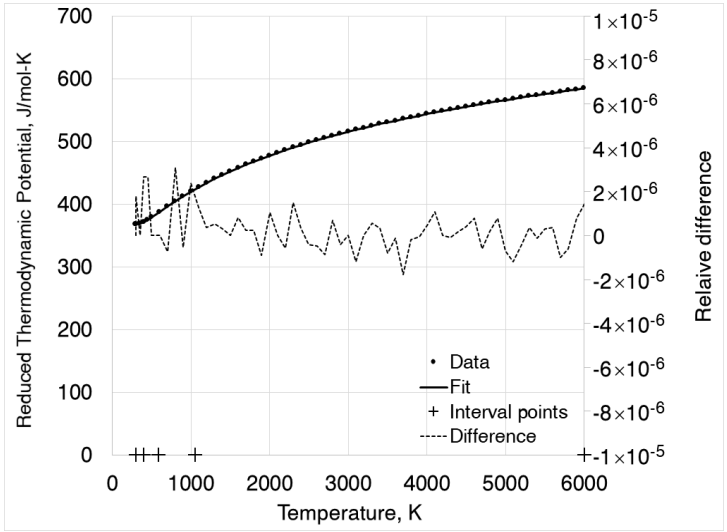
Zirconium Chloride ($ZrCl_4$) – Cl4Zr1(g)

MW = 233.03 g/mol – $\Delta_f H^o$(298.15 K) = -869.979 kJ/mol – S^o(298.15 K) = 367.716 J/mol-K – p^o = 0.1 MPa



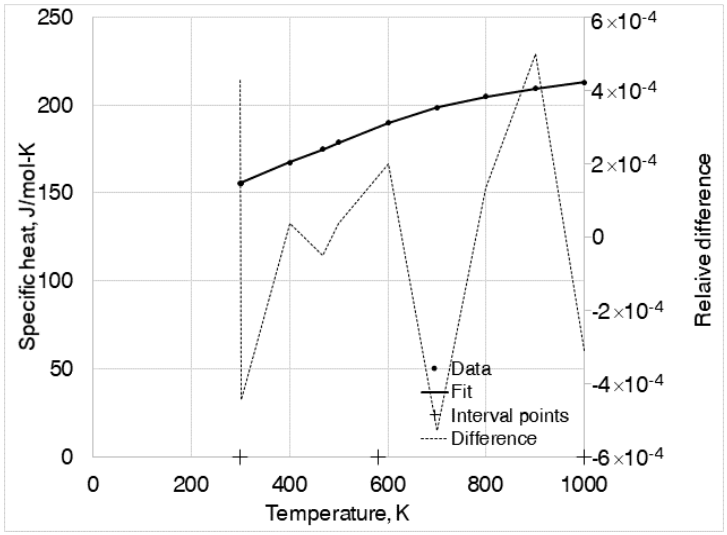


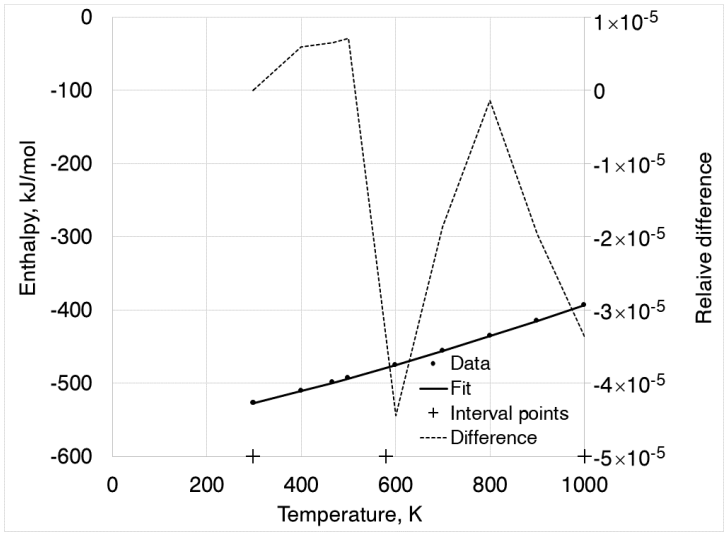


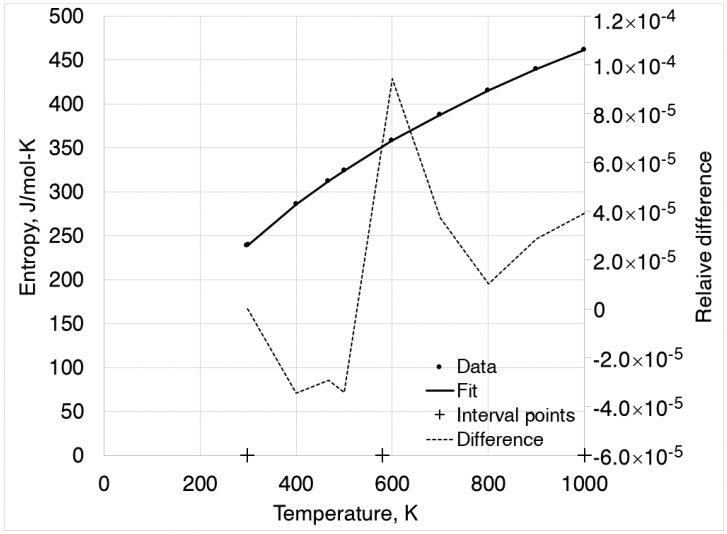


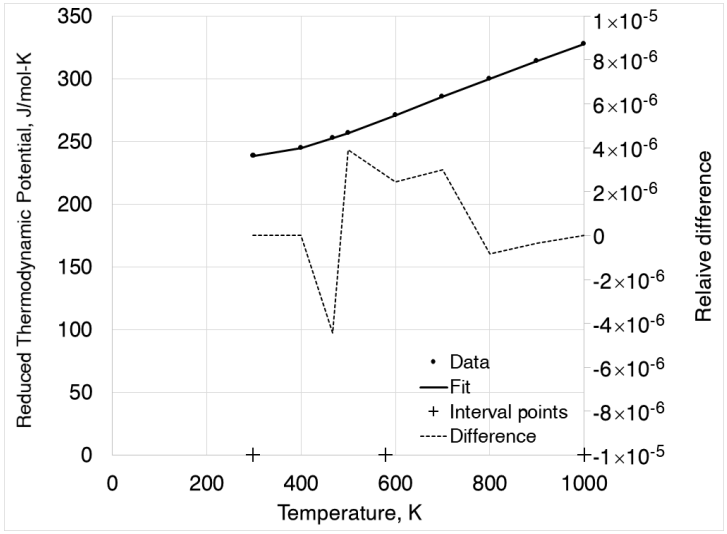
Molybdenum Chloride ($MoCl_5$) – Cl5Mo1(cr)

MW = 273.21 g/mol – $\Delta_f H^o$(298.15 K) = -527.184 kJ/mol – S^o(298.15 K) = 238.488 J/mol-K – p^o = 0.1 MPa



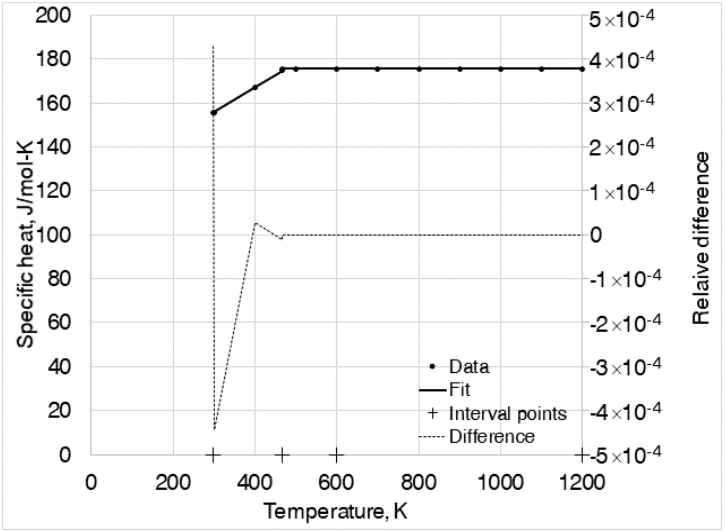


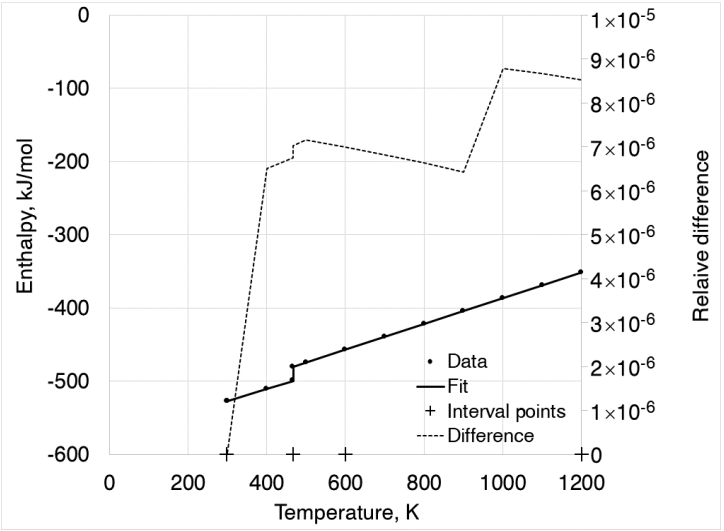


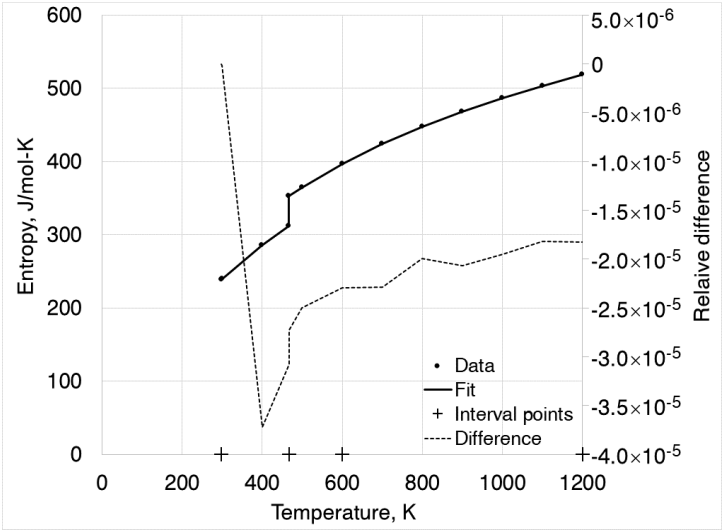


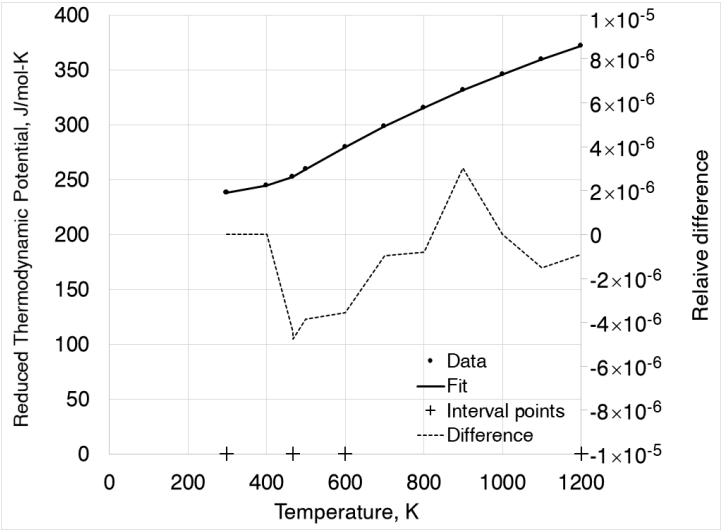
Molybdenum Chloride ($MoCl_5$) – Cl5Mo1(cr,l)

MW = 273.21 g/mol – $\Delta_f H^o$(298.15 K) = -527.184 kJ/mol – S^o(298.15 K) = 238.488 J/mol-K – p^o = 0.1 MPa



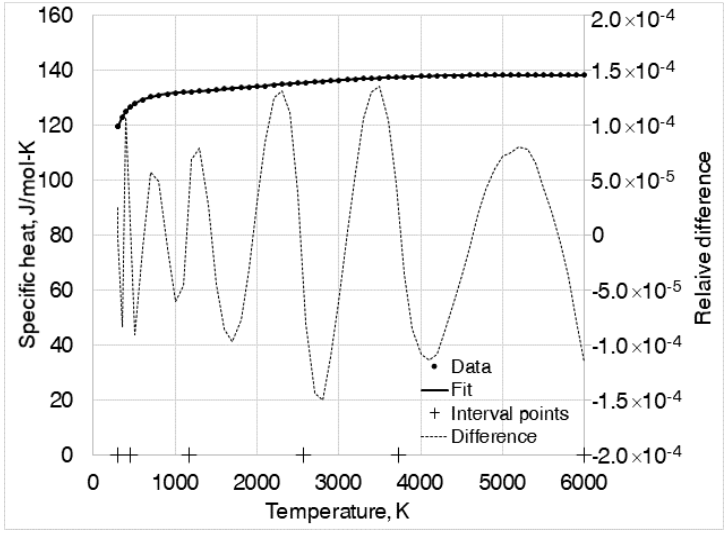


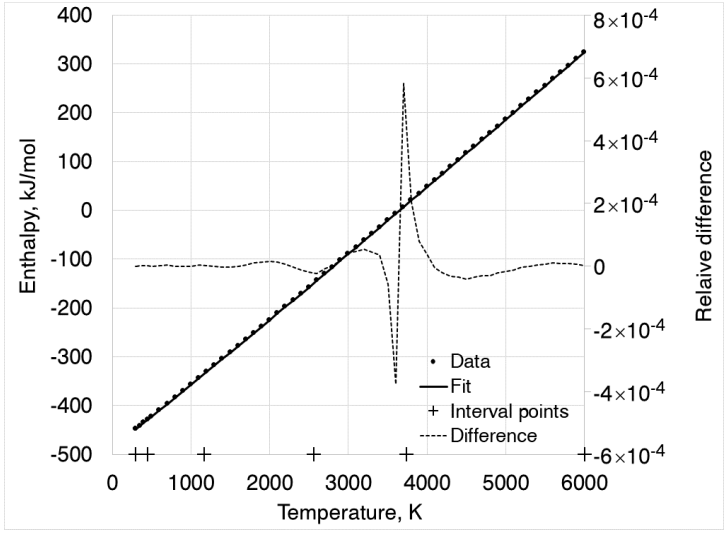


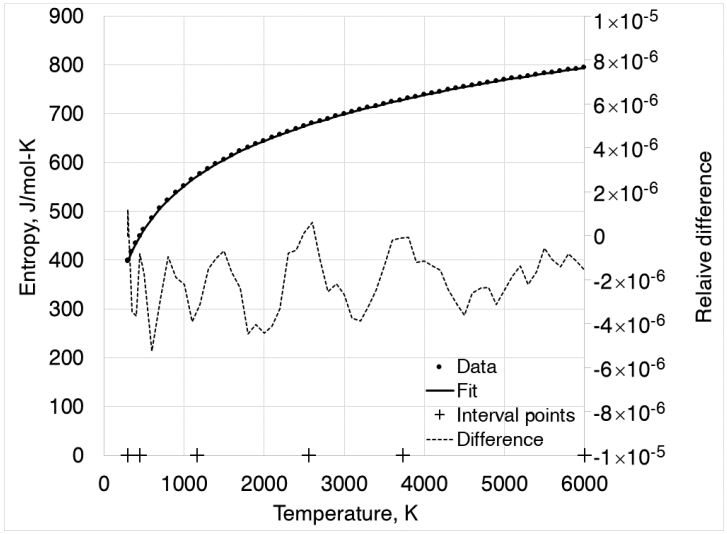


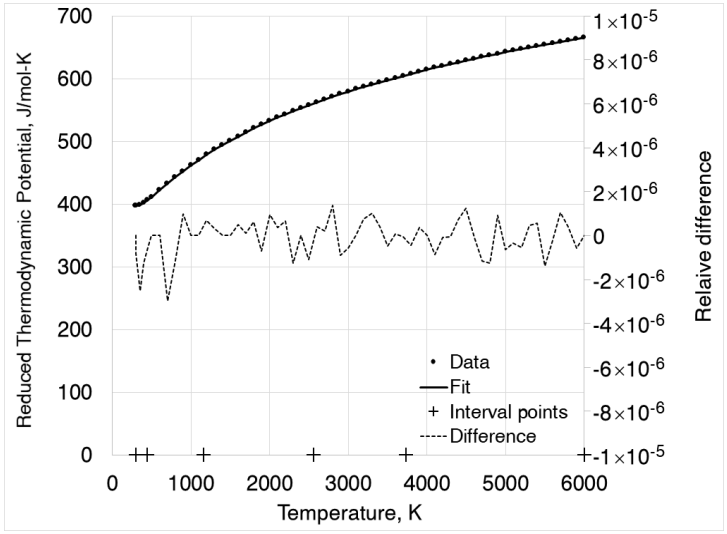
Molybdenum Chloride ($MoCl_5$) – Cl5Mo1(g)

MW = 273.21 g/mol – $\Delta_f H^o$(298.15 K) = -447.688 kJ/mol – S^o(298.15 K) = 397.774 J/mol-K – p^o = 0.1 MPa



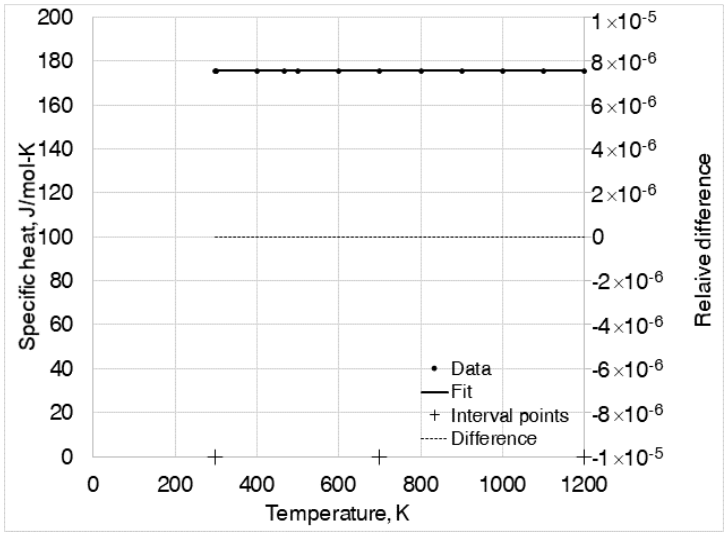


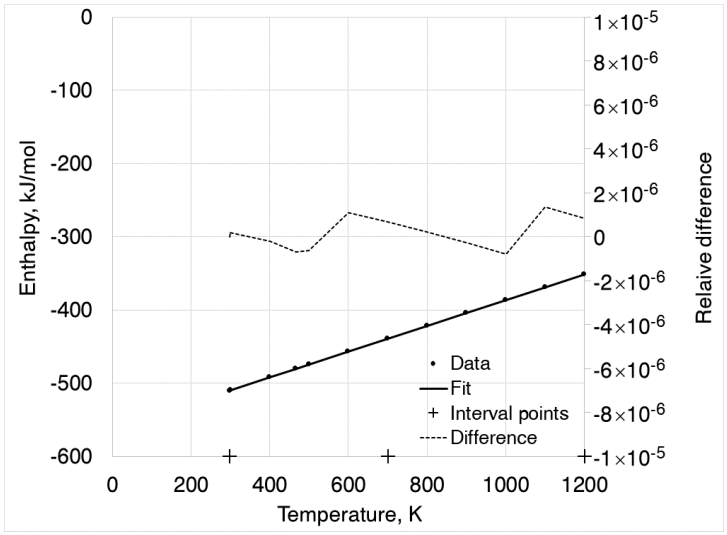


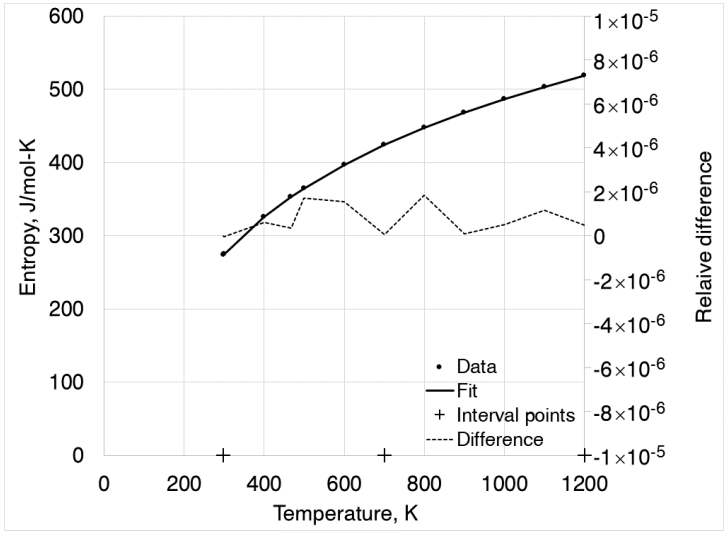


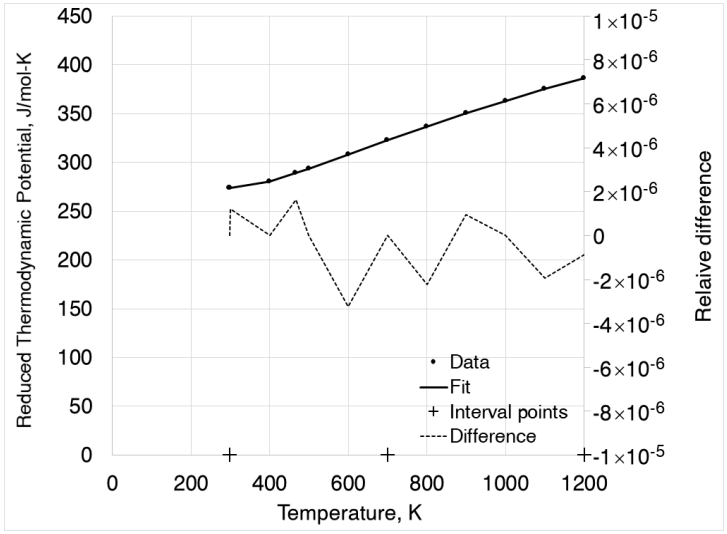
Molybdenum Chloride ($MoCl_5$) – Cl5Mo1(l)

MW = 273.21 g/mol – $\Delta_f H^o$(298.15 K) = -510.114 kJ/mol – S^o(298.15 K) = 273.807 J/mol-K – p^o = 0.1 MPa



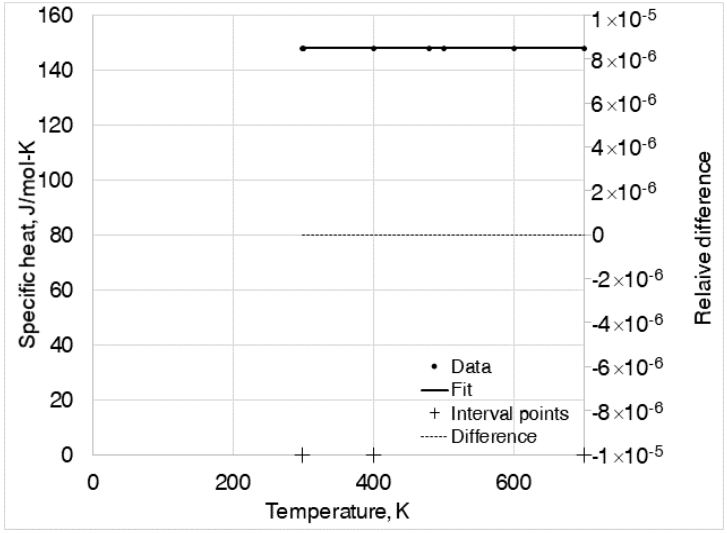


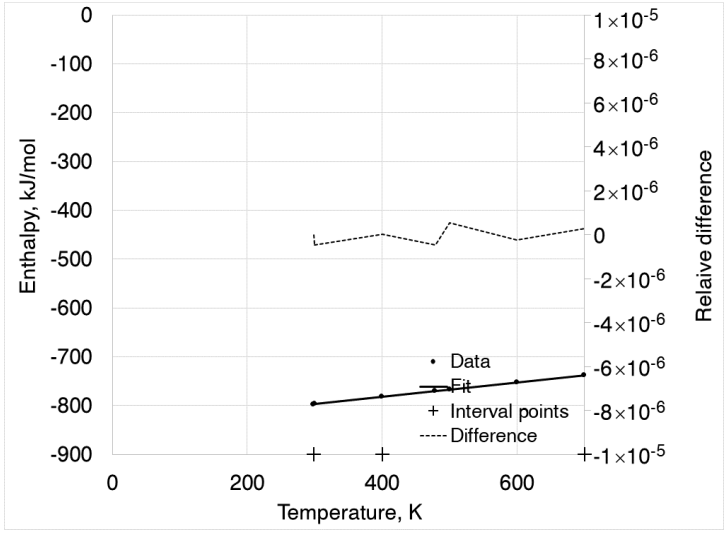


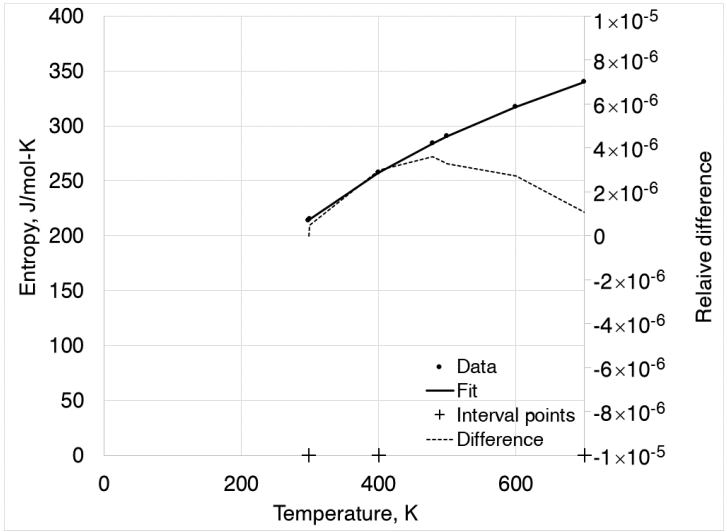


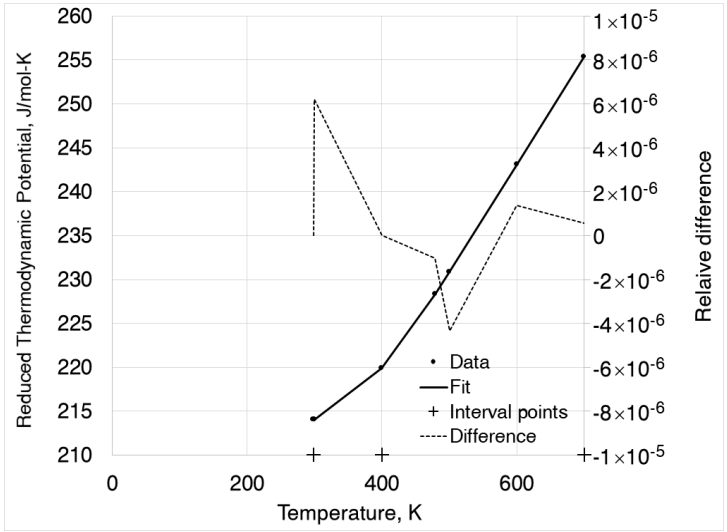
Niobium Chloride ($NbCl_5$) – Cl5Nb1(cr)

MW = 270.171 g/mol – $\Delta_f H^o$(298.15 K) = -797.47 kJ/mol – S^o(298.15 K) = 214.053 J/mol-K – p^o = 0.1 MPa



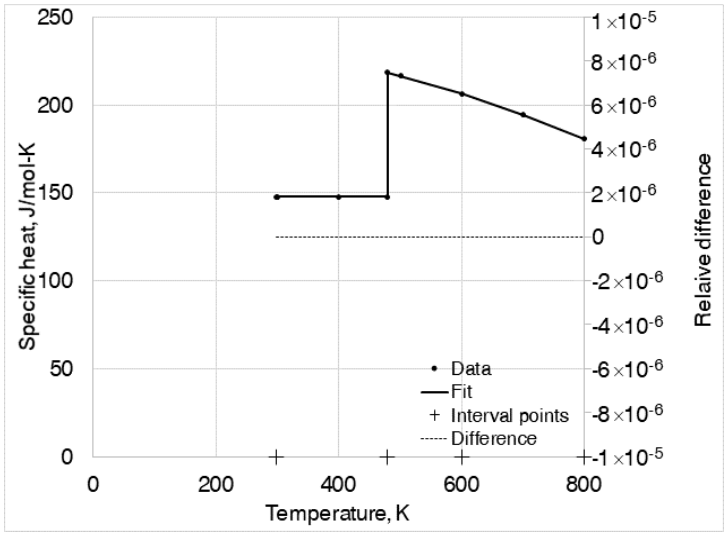


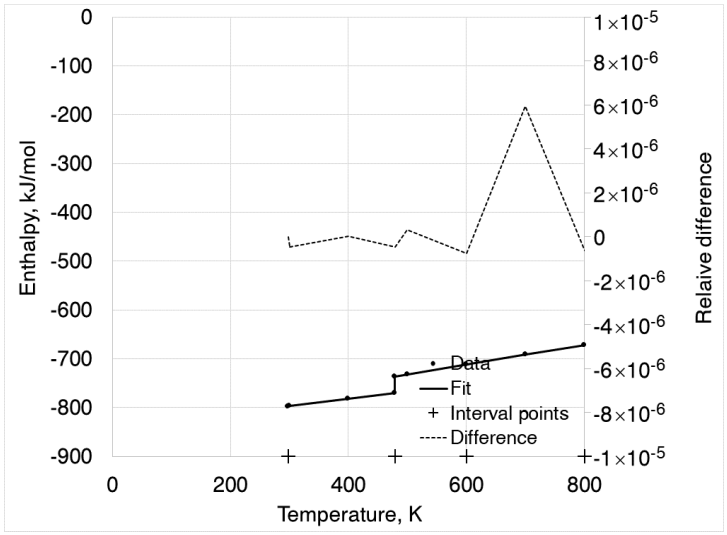


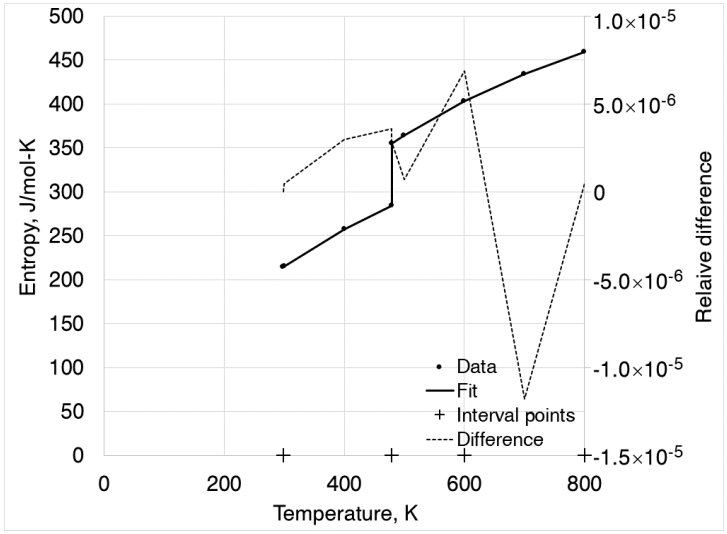


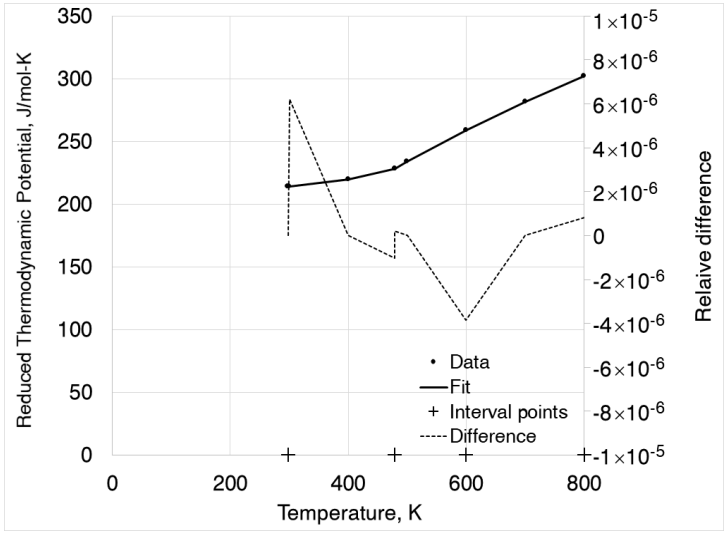
Niobium Chloride ($NbCl_5$) – Cl5Nb1(cr,l)

MW = 270.171 g/mol – $\Delta_f H^o$(298.15 K) = -797.47 kJ/mol – S^o(298.15 K) = 214.053 J/mol-K – p^o = 0.1 MPa



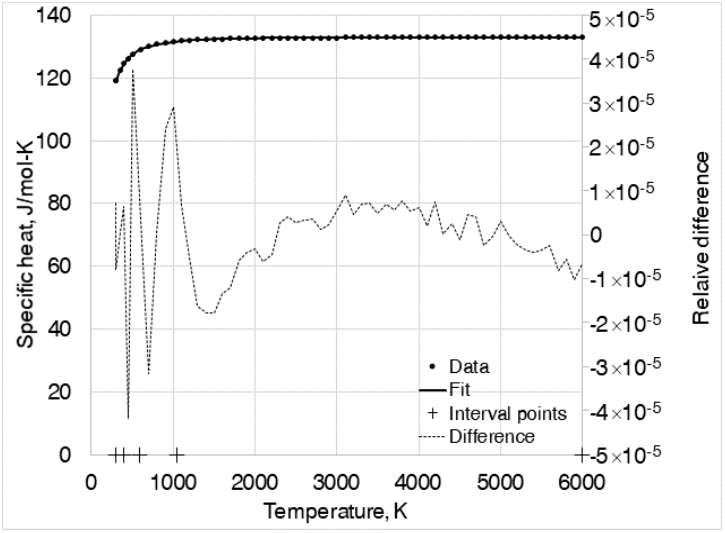


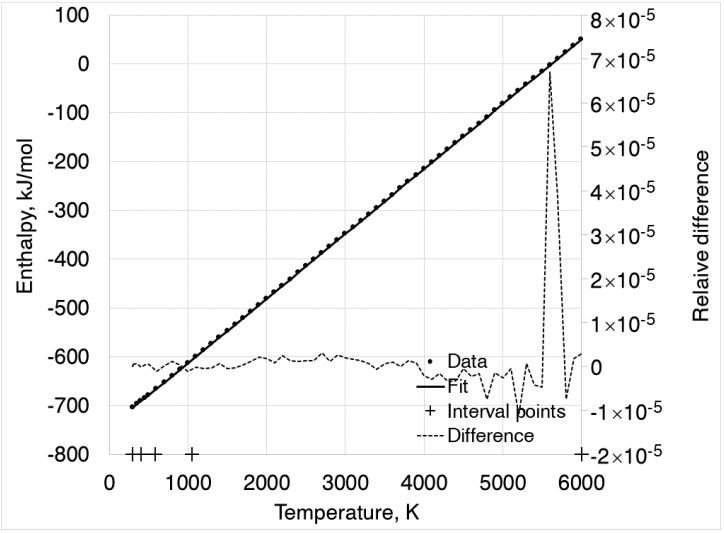


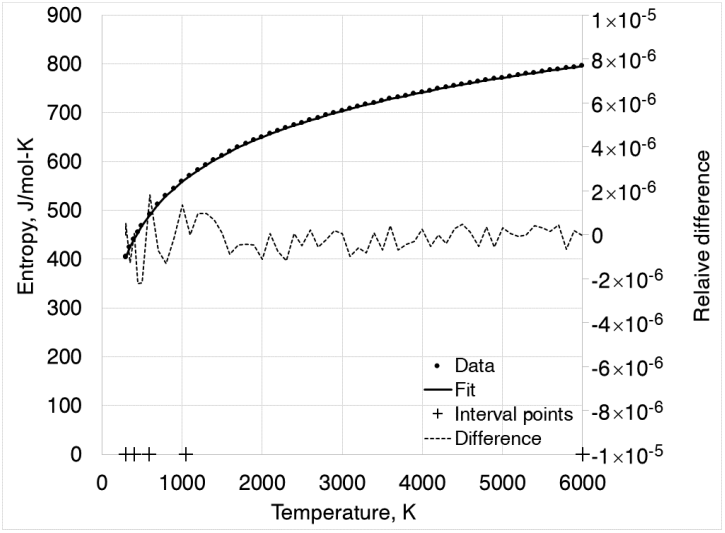


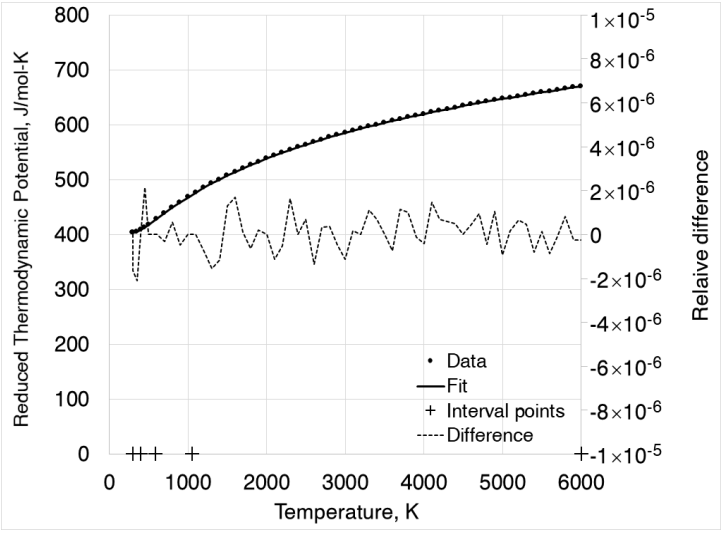
Niobium Chloride ($NbCl_5$) – Cl5Nb1(g)

MW = 270.171 g/mol – $\Delta_f H^o$(298.15 K) = -703.33 kJ/mol – S^o(298.15 K) = 404.111 J/mol-K – p^o = 0.1 MPa



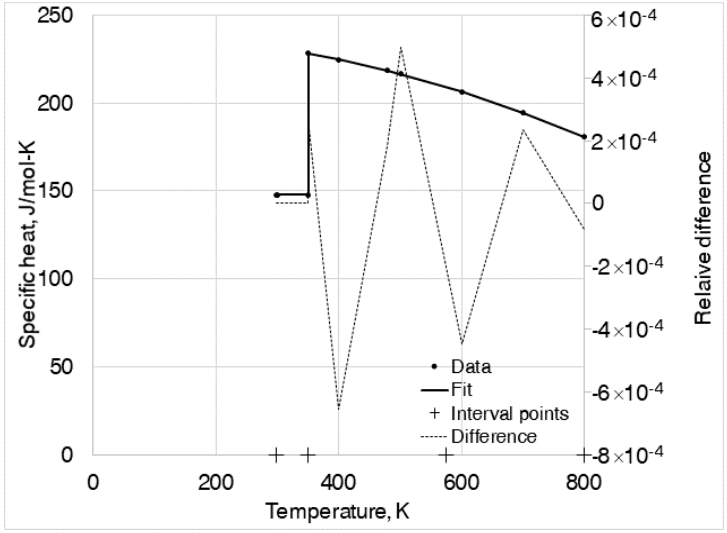


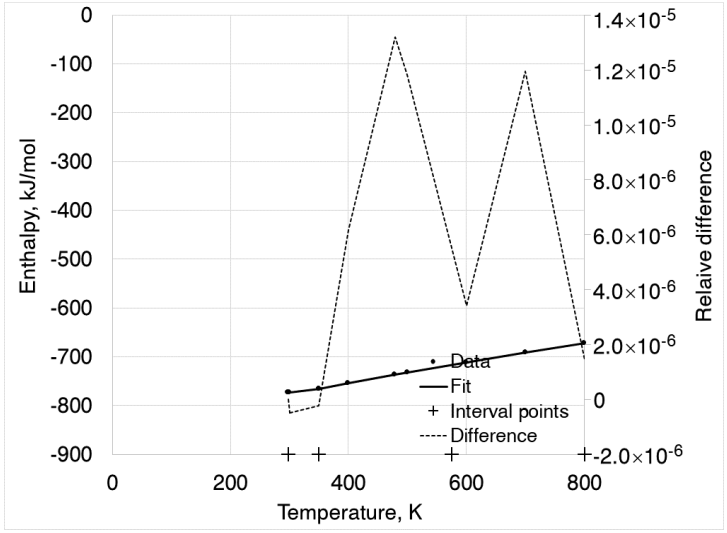


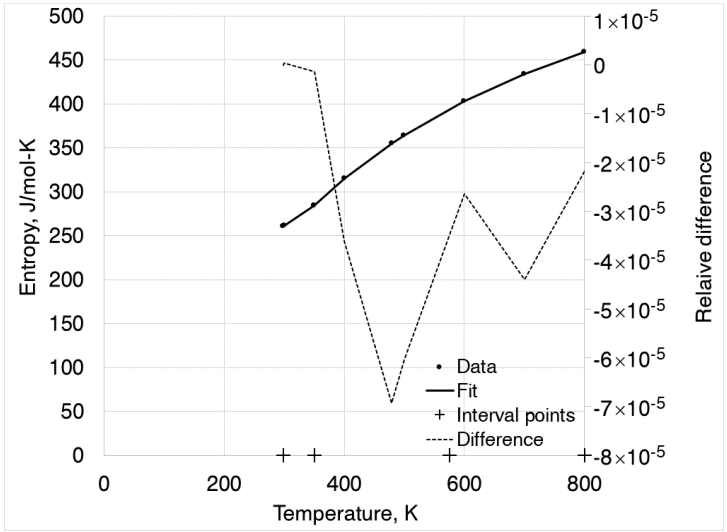


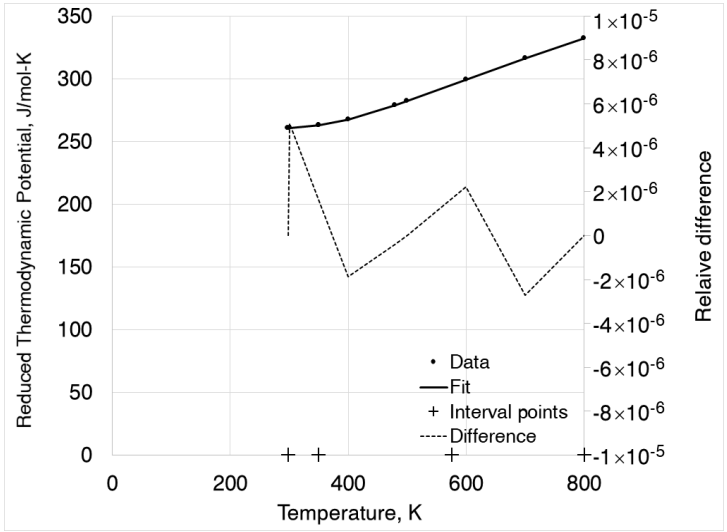
Niobium Chloride ($NbCl_5$) – Cl5Nb1(l)

MW = 270.171 g/mol – $\Delta_f H^o$(298.15 K) = -773.354 kJ/mol – S^o(298.15 K) = 260.97 J/mol-K – p^o = 0.1 MPa



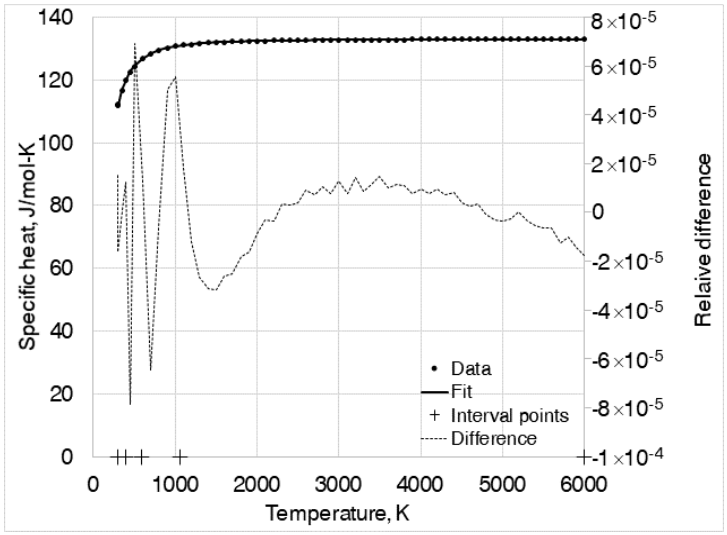


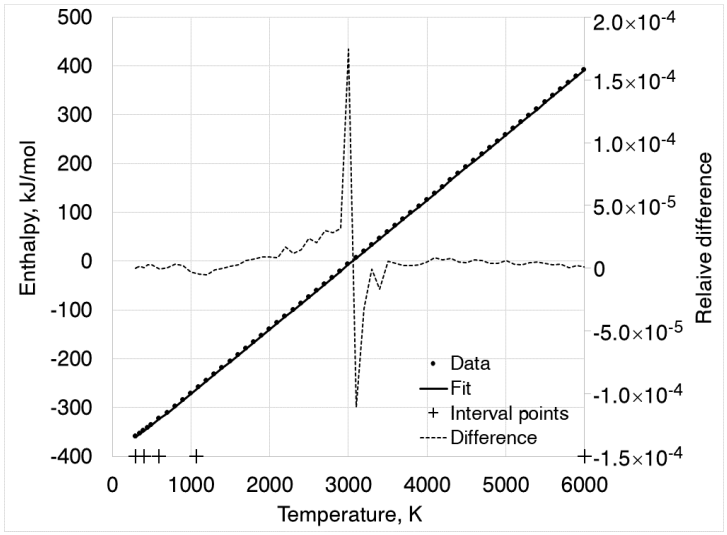


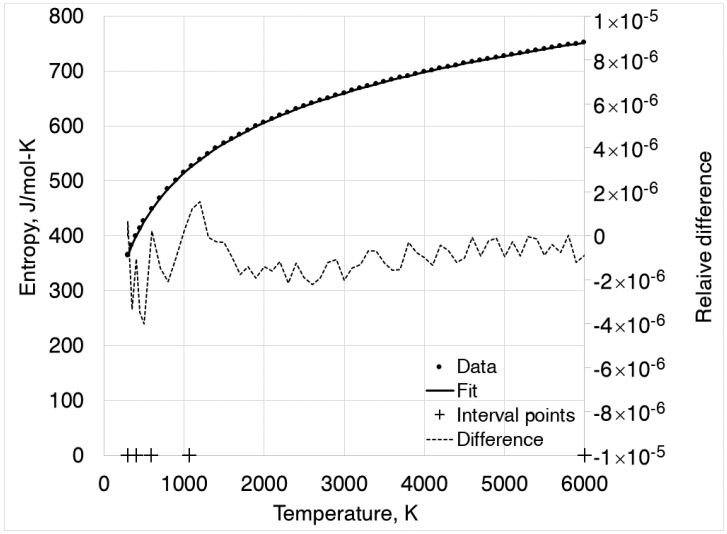


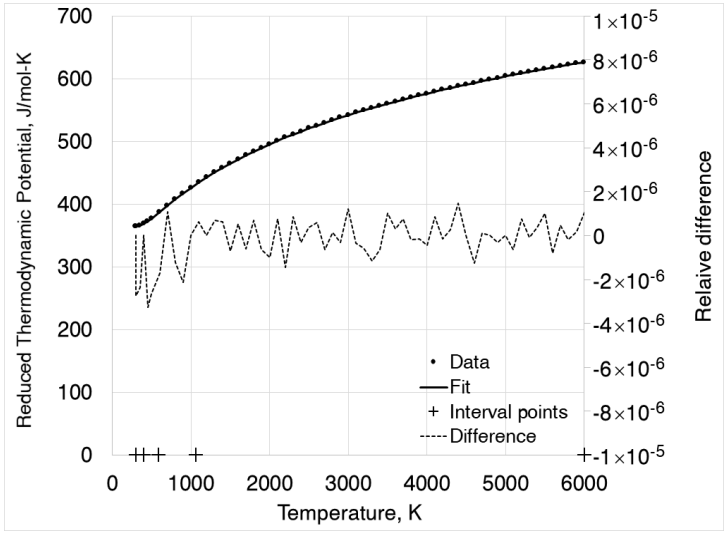
Phosphorus Chloride (PCl_5) – Cl5P1(g)

MW = 208.239 g/mol – $\Delta_f H^o$(298.15 K) = -360.184 kJ/mol – S^o(298.15 K) = 364.288 J/mol-K – p^o = 0.1 MPa



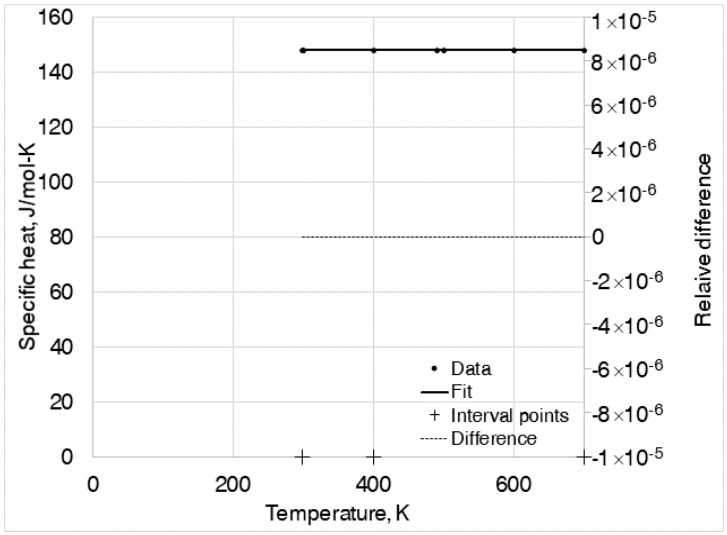


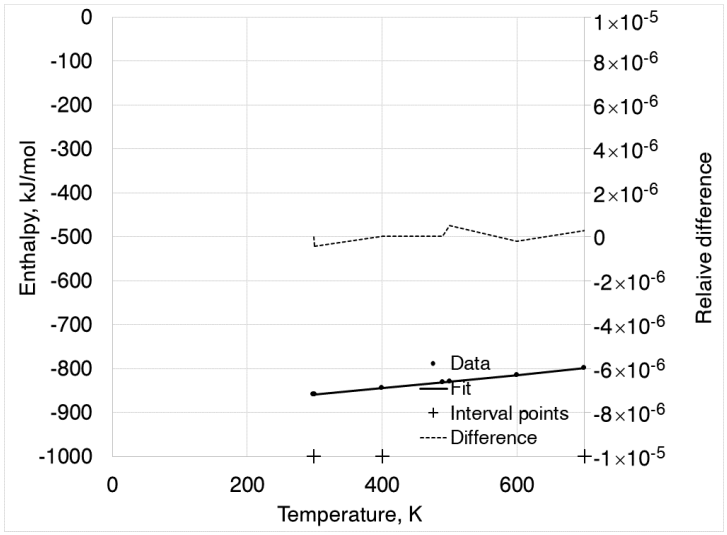


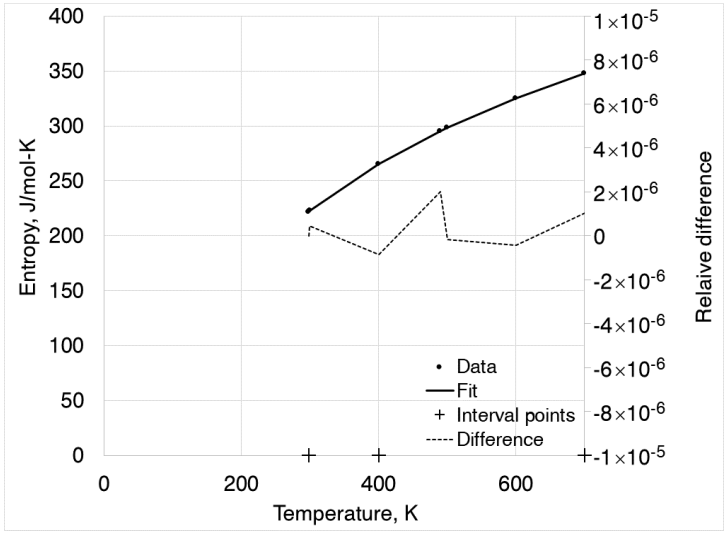


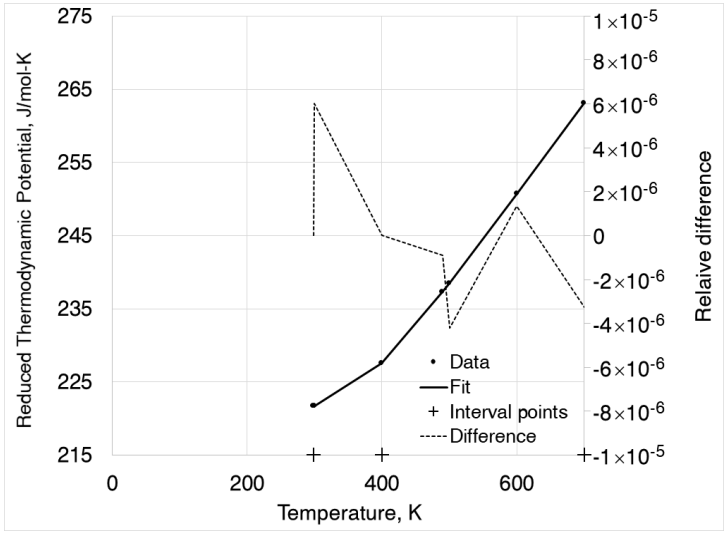
Tantalum Chloride ($TaCl_5$) – Cl5Ta1(cr)

MW = 358.213 g/mol – $\Delta_f H^o$(298.15 K) = -858.975 kJ/mol – S^o(298.15 K) = 221.752 J/mol-K – p^o = 0.1 MPa



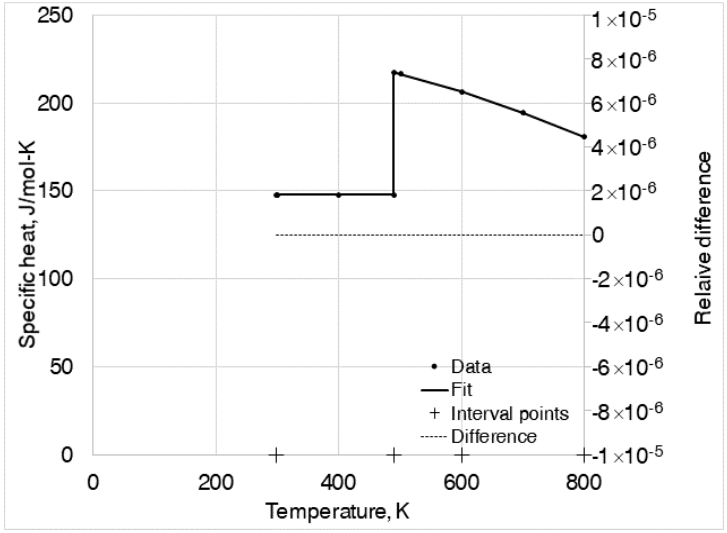


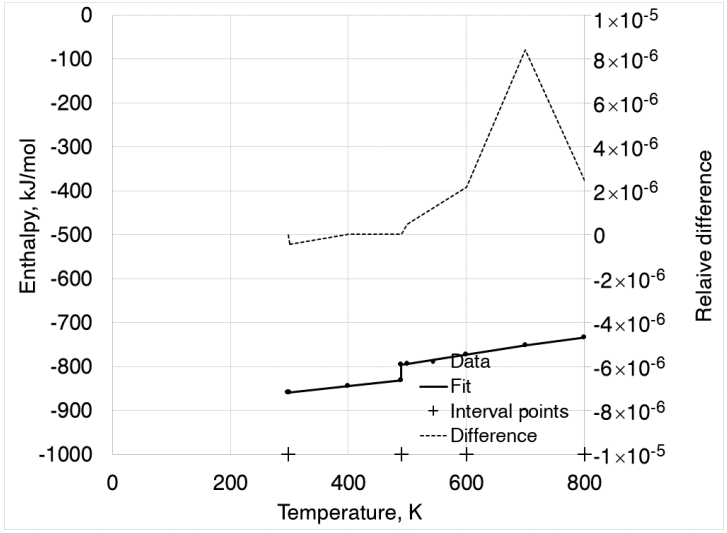


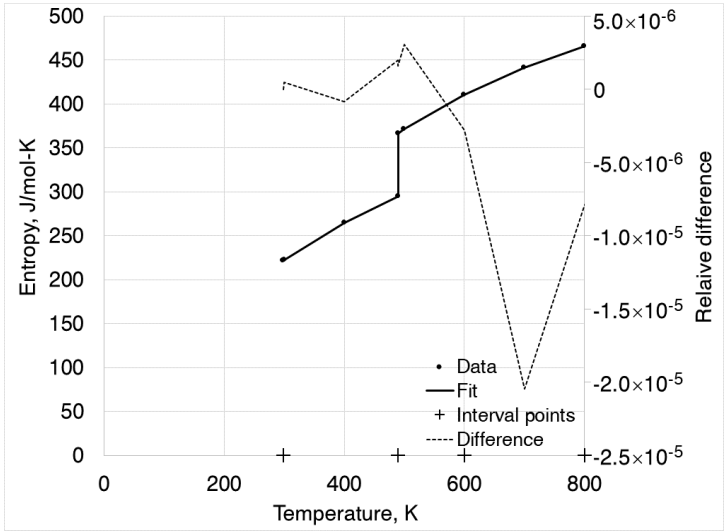


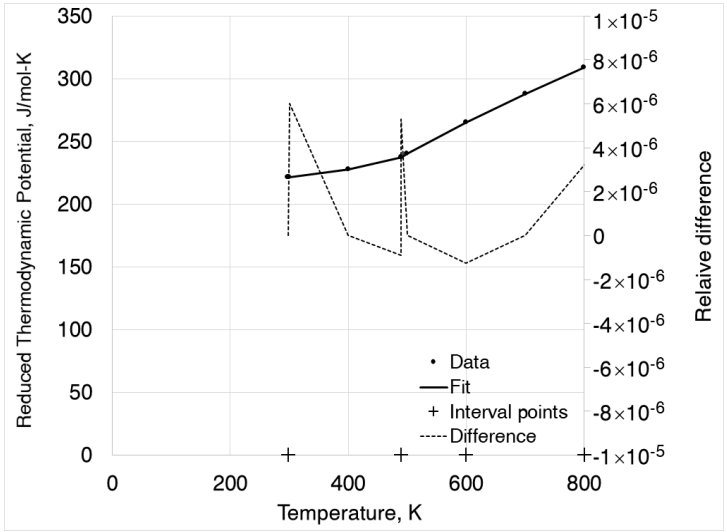
Tantalum Chloride ($TaCl_5$) – Cl5Ta1(cr,l)

MW = 358.213 g/mol – $\Delta_f H^o$(298.15 K) = -858.975 kJ/mol – S^o(298.15 K) = 221.752 J/mol-K – p^o = 0.1 MPa



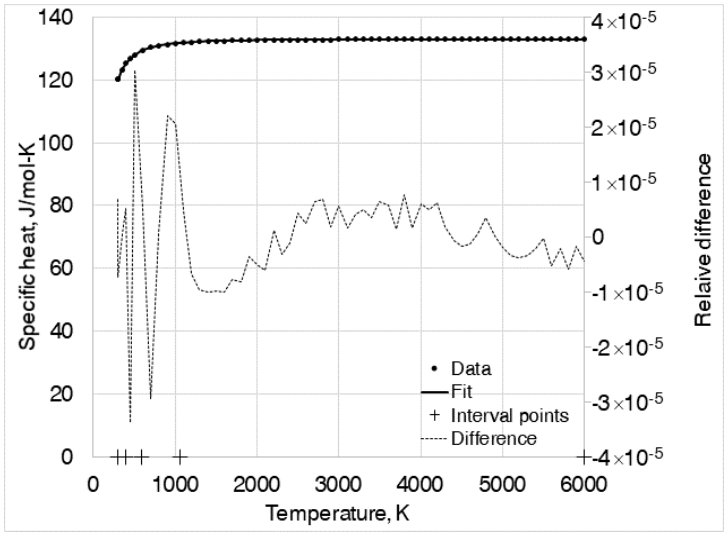


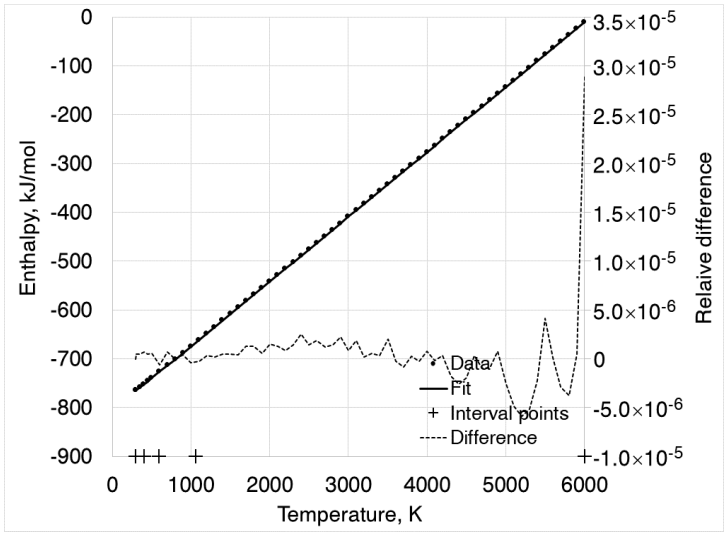


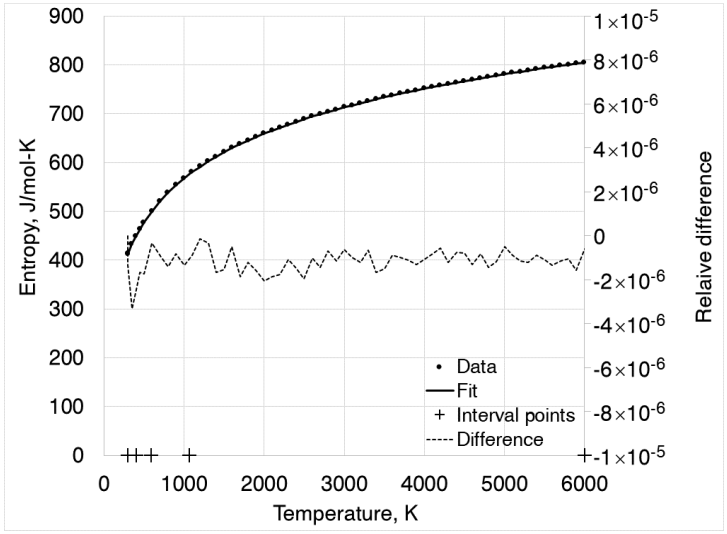


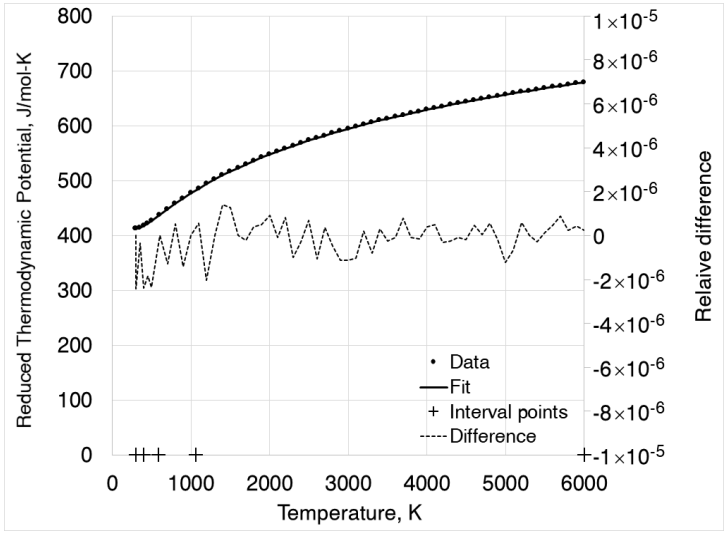
Tantalum Chloride ($TaCl_5$) – Cl5Ta1(g)

MW = 358.213 g/mol – $\Delta_f H^o$(298.15 K) = -764.835 kJ/mol – S^o(298.15 K) = 412.973 J/mol-K – p^o = 0.1 MPa



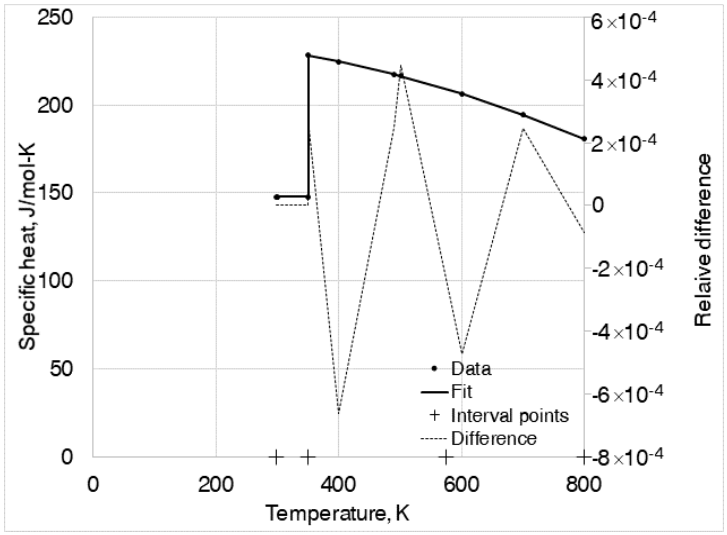


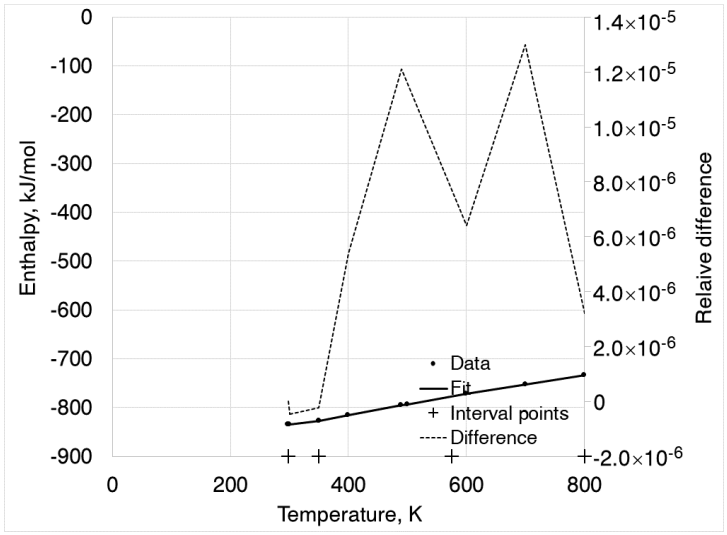


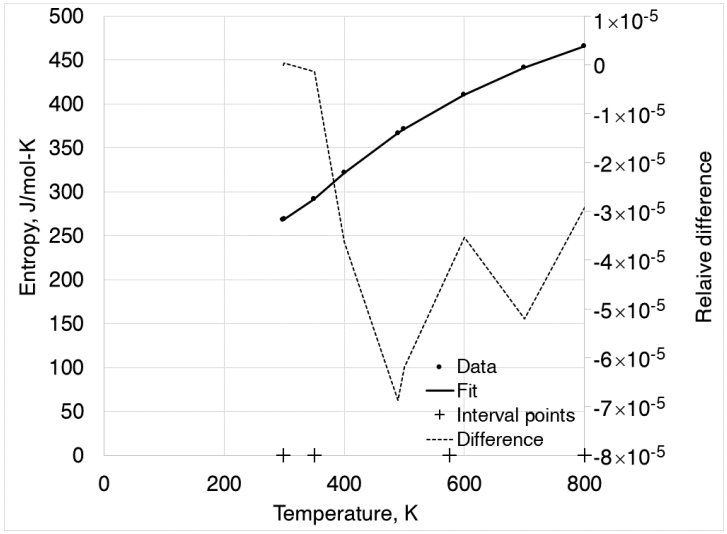


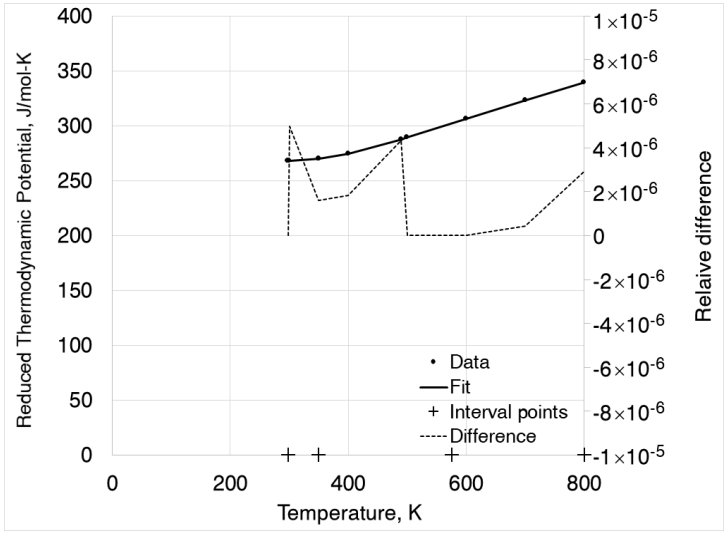
Tantalum Chloride ($TaCl_5$) – Cl5Ta1(l)

MW = 358.213 g/mol – $\Delta_f H^o$(298.15 K) = -834.347 kJ/mol – S^o(298.15 K) = 268.135 J/mol-K – p^o = 0.1 MPa



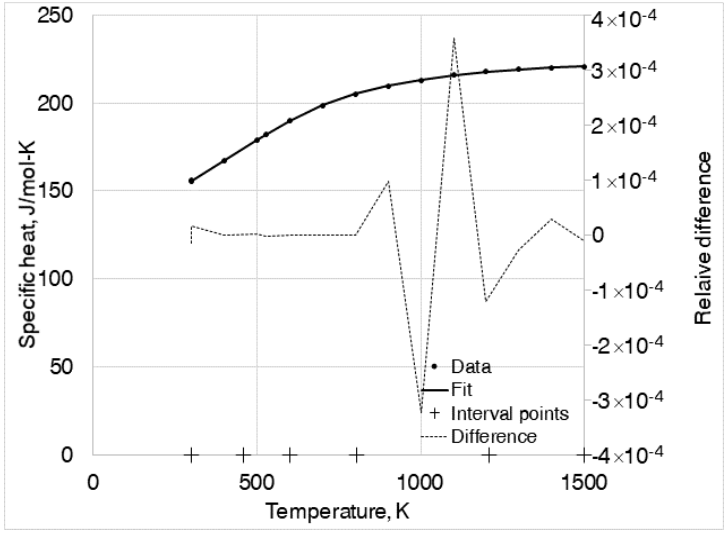


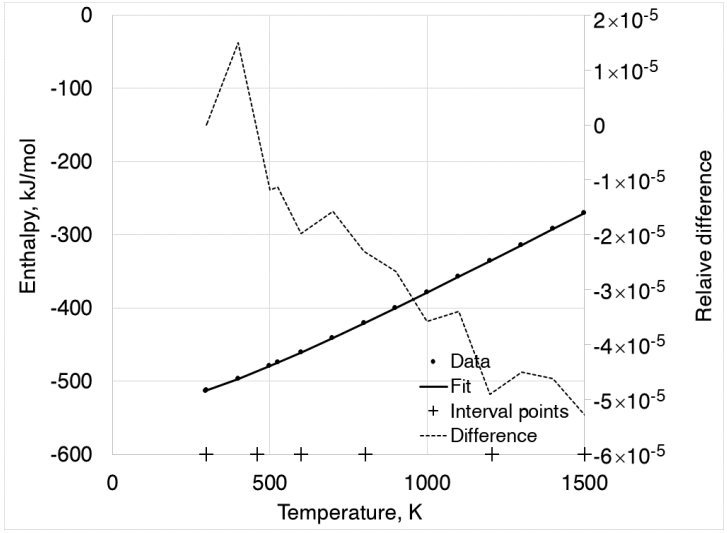


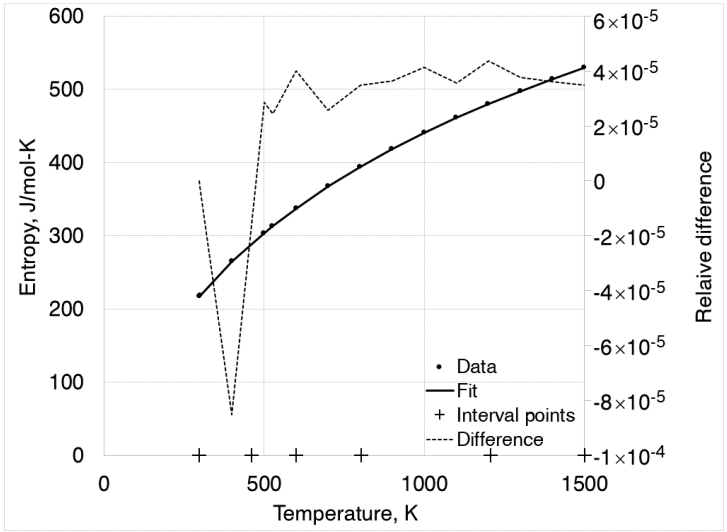


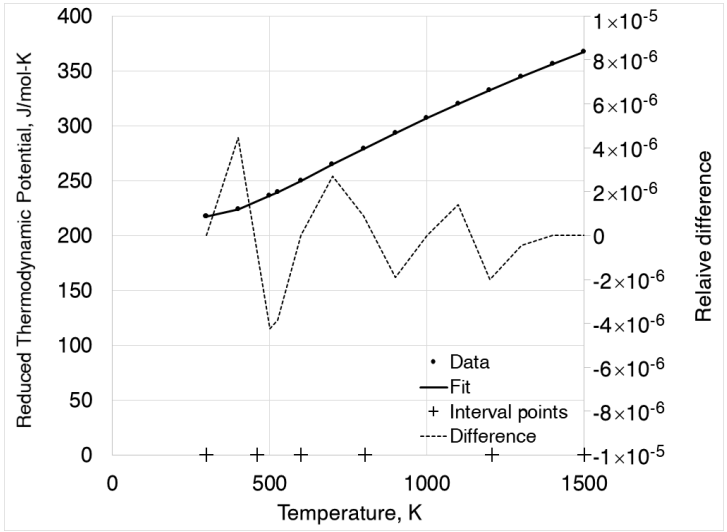
Tungsten Chloride (WCl_5) – Cl5W1(cr)

MW = 361.12 g/mol – $\Delta_f H^o$(298.15 K) = -512.958 kJ/mol – S^o(298.15 K) = 217.568 J/mol-K – p^o = 0.1 MPa



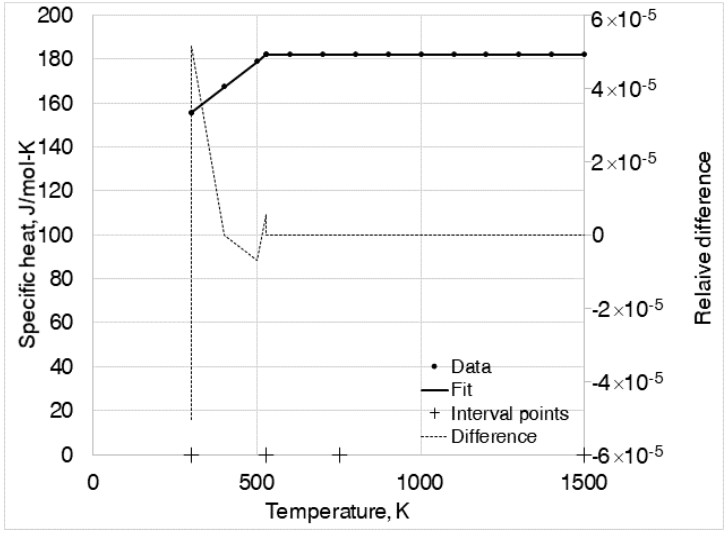


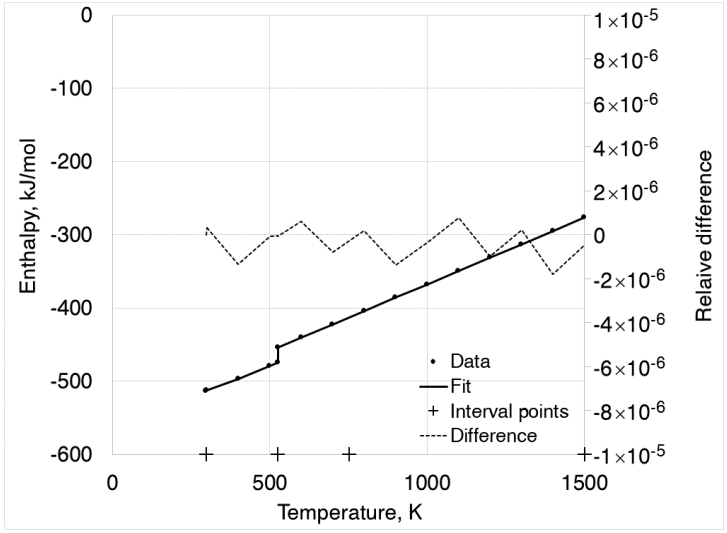


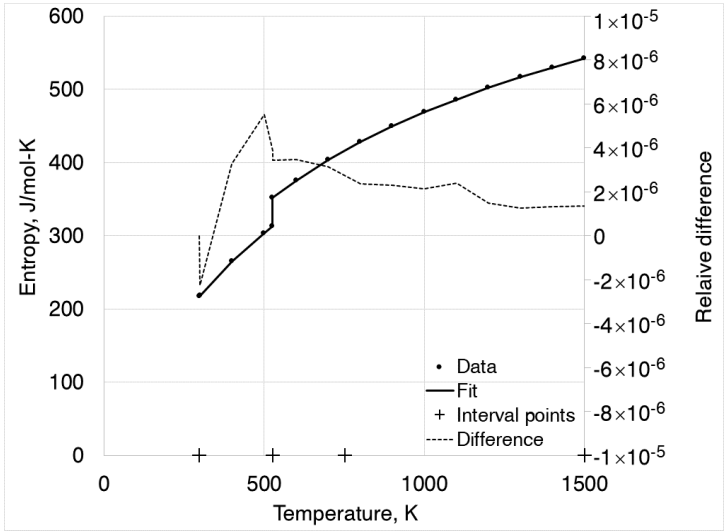


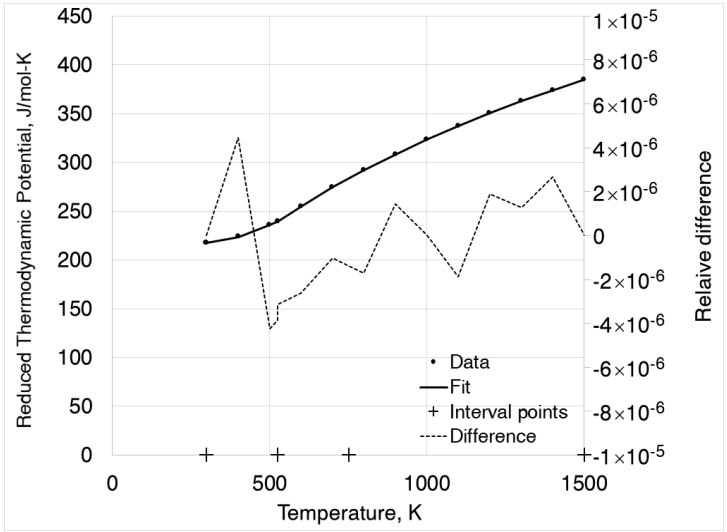
Tungsten Chloride (WCl_5) – Cl5W1(cr,l)

MW = 361.12 g/mol – $\Delta_f H^o$(298.15 K) = -512.958 kJ/mol – S^o(298.15 K) = 217.568 J/mol-K – p^o = 0.1 MPa



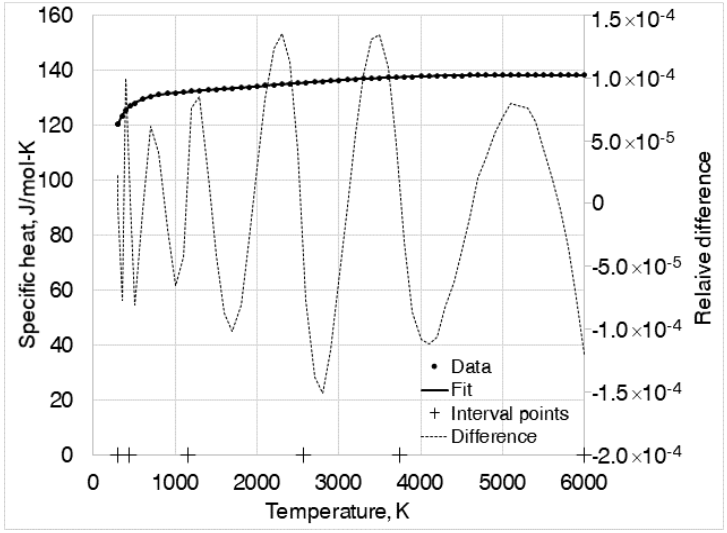


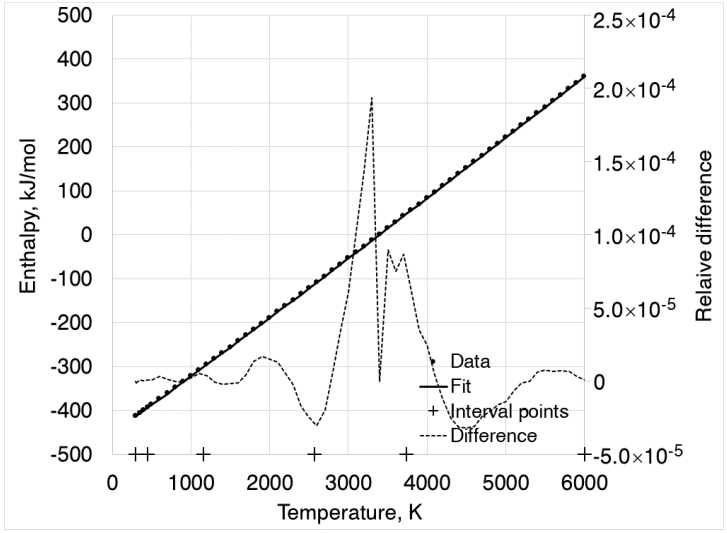


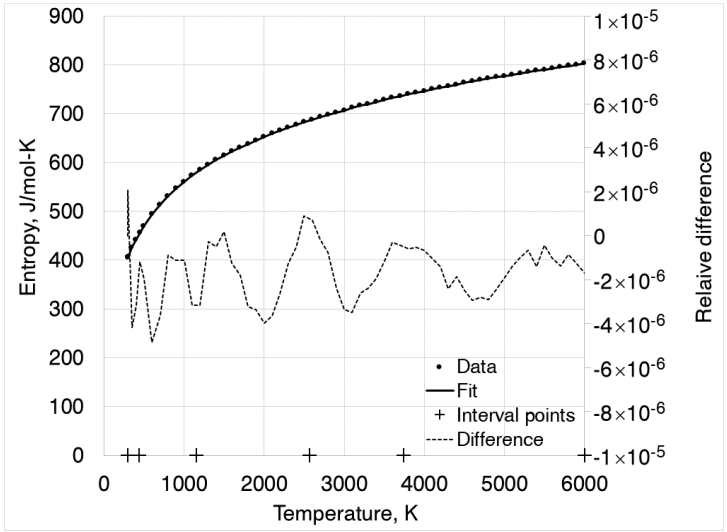


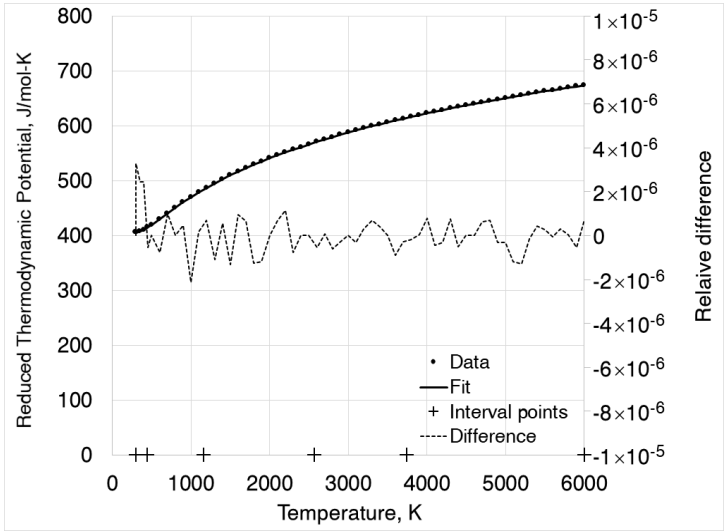
Tungsten Chloride (WCl_5) – Cl5W1(g)

MW = 361.12 g/mol – $\Delta_f H^o$(298.15 K) = -412.542 kJ/mol – S^o(298.15 K) = 405.693 J/mol-K – p^o = 0.1 MPa



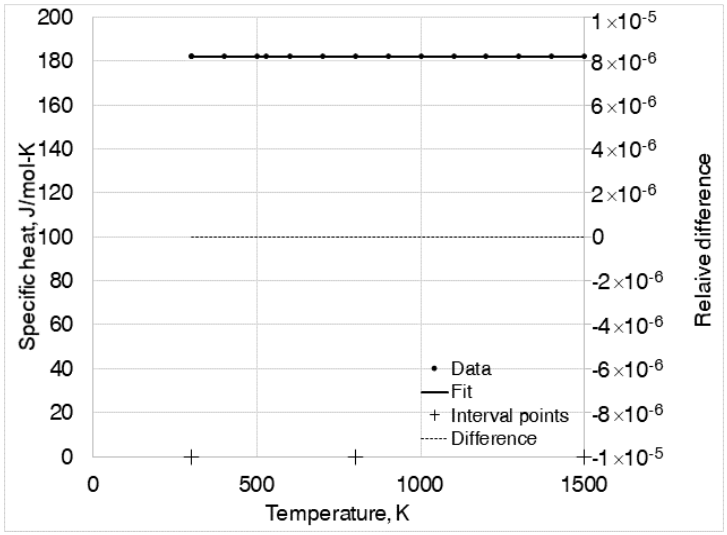


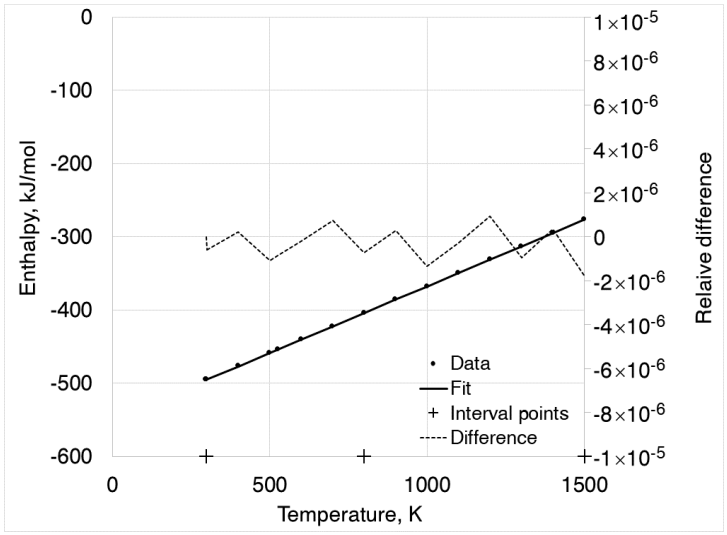


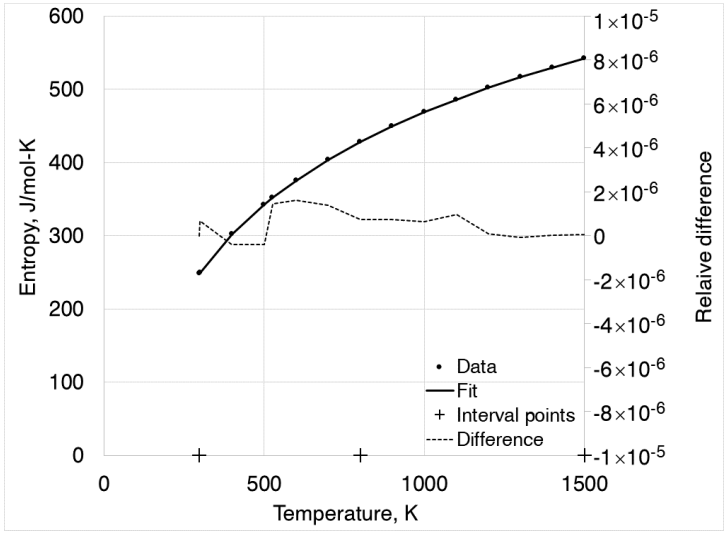


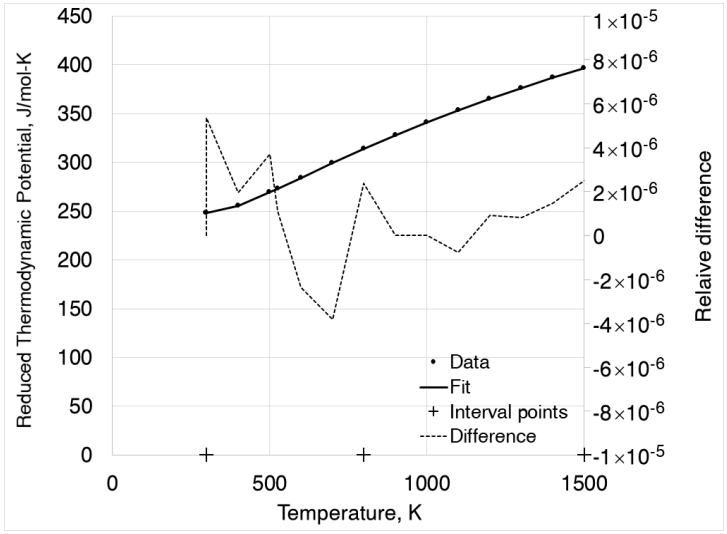
Tungsten Chloride (WCl_5) – Cl5W1(l)

MW = 361.12 g/mol – $\Delta_f H^o$(298.15 K) = -495.387 kJ/mol – S^o(298.15 K) = 248.498 J/mol-K – p^o = 0.1 MPa



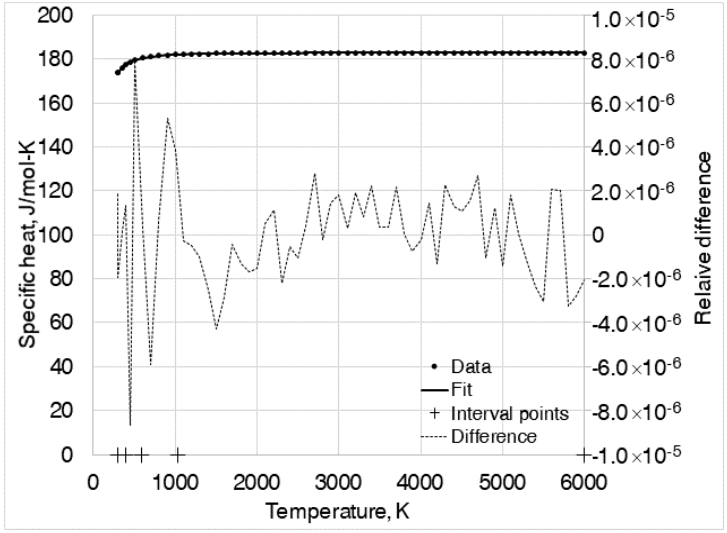


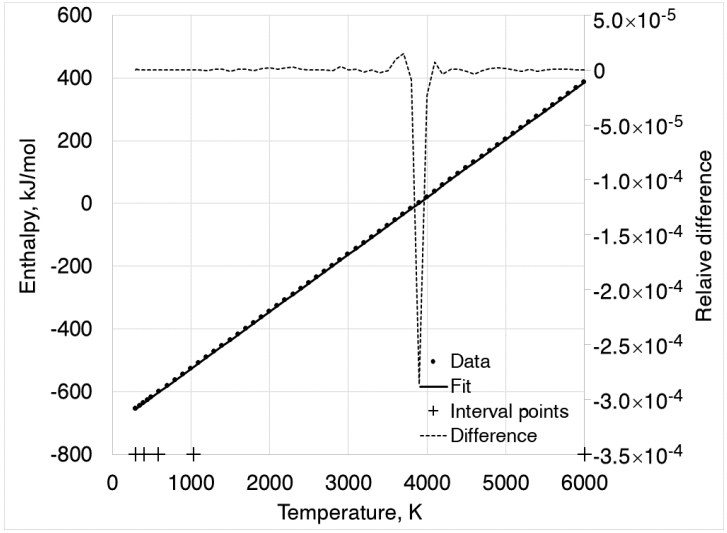


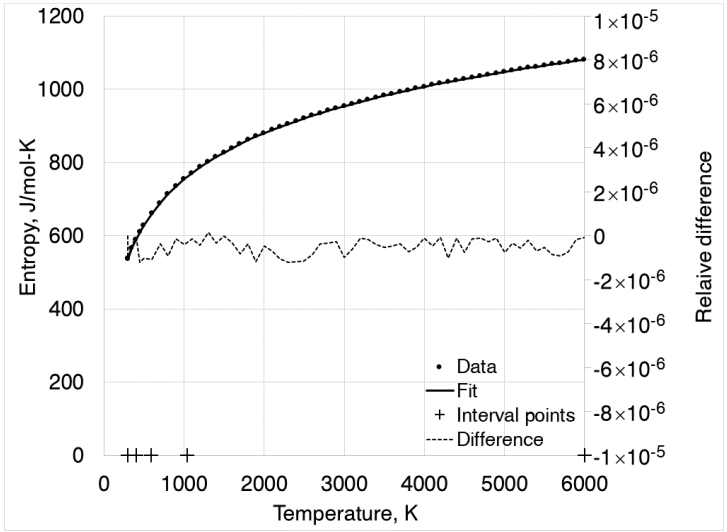


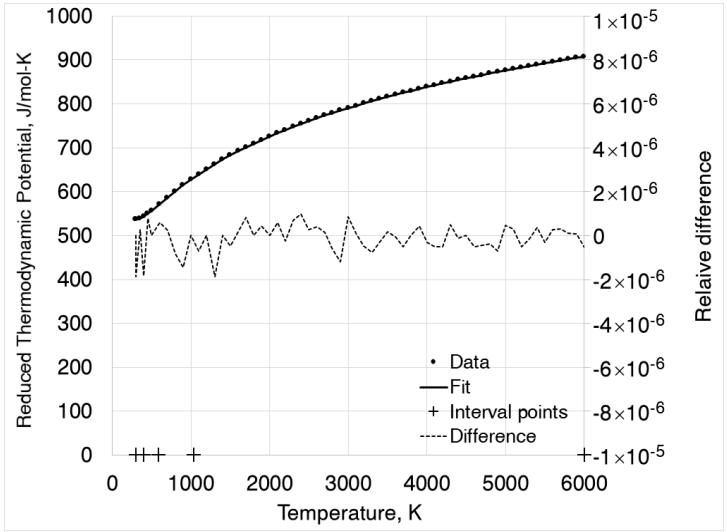
Iron Chloride ($(FeCl_3)_2$) – Cl6Fe2(g)

MW = 324.412 g/mol – $\Delta_f H^o$(298.15 K) = -654.378 kJ/mol – S^o(298.15 K) = 536.937 J/mol-K – p^o = 0.1 MPa



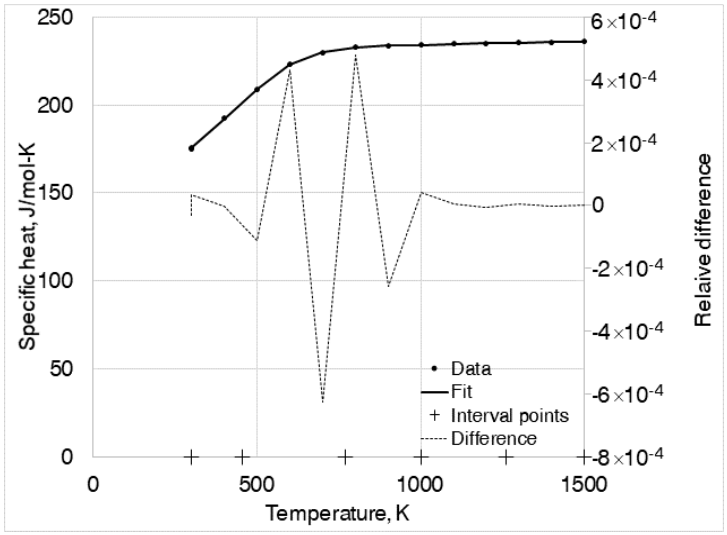


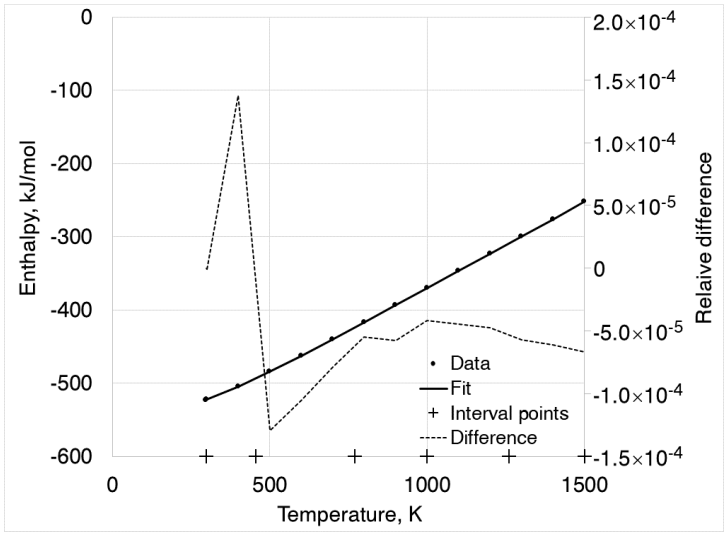


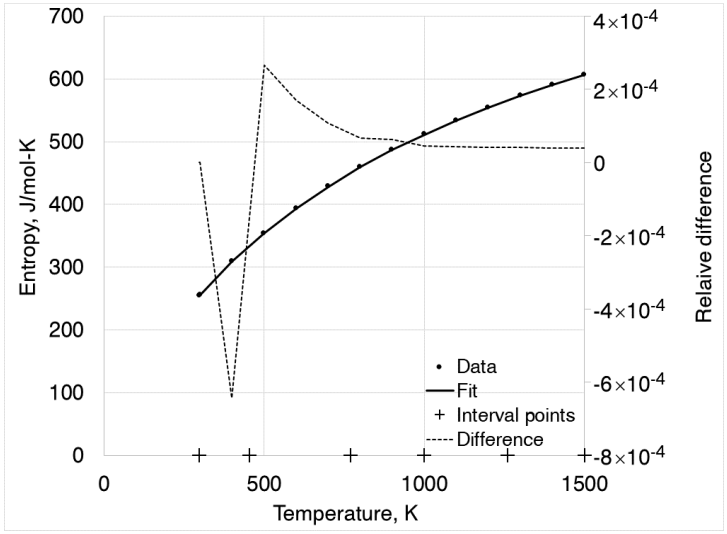


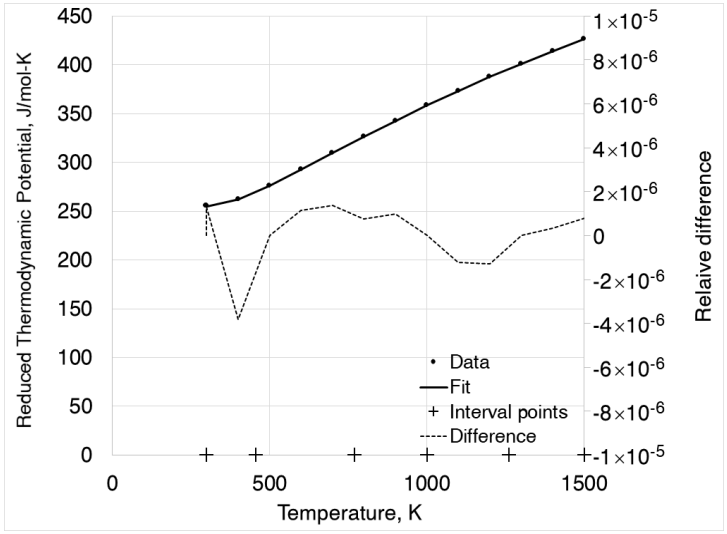
Molybdenum Chloride ($MoCl_6$) – Cl6Mo1(cr)

MW = 308.66 g/mol – $\Delta_f H^o$(298.15 K) = -523 kJ/mol – S^o(298.15 K) = 255.224 J/mol-K – p^o = 0.1 MPa



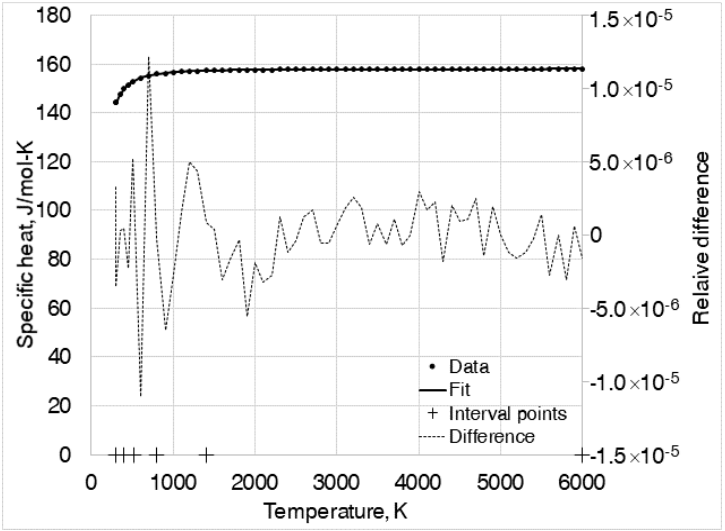


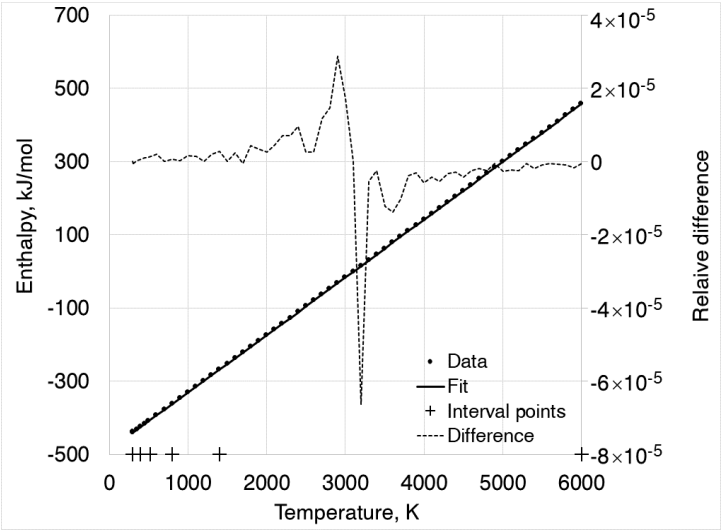


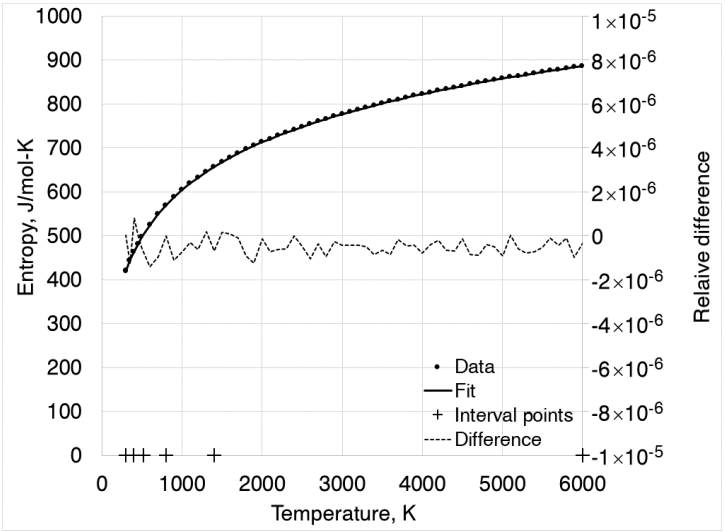


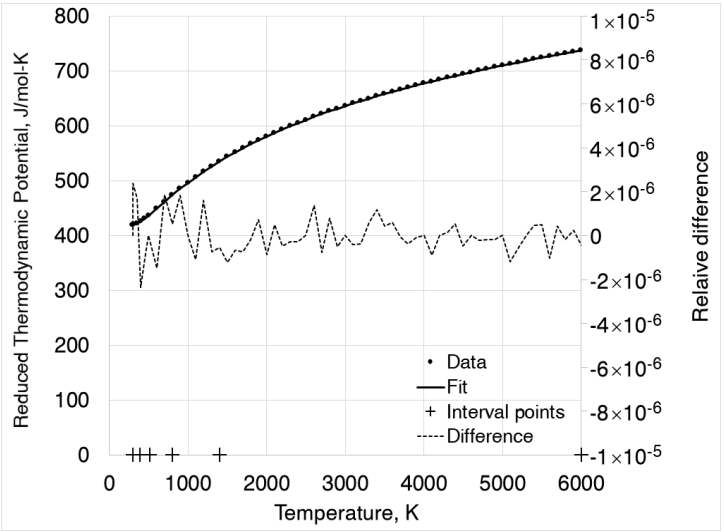
Molybdenum Chloride ($MoCl_6$) – Cl6Mo1(g)

MW = 308.66 g/mol – $\Delta_f H^o$(298.15 K) = -439.32 kJ/mol – S^o(298.15 K) = 419.654 J/mol-K – p^o = 0.1 MPa



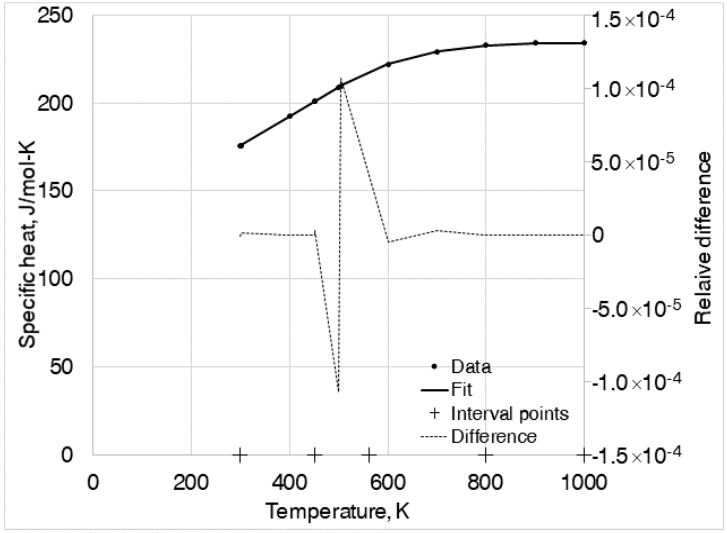


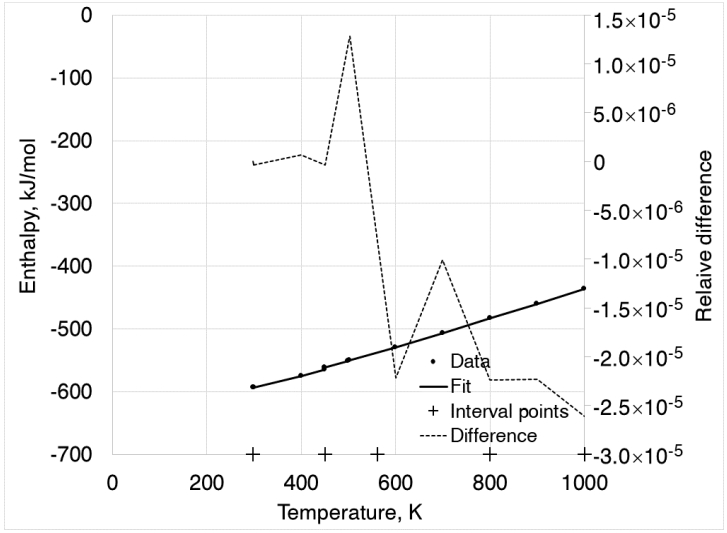


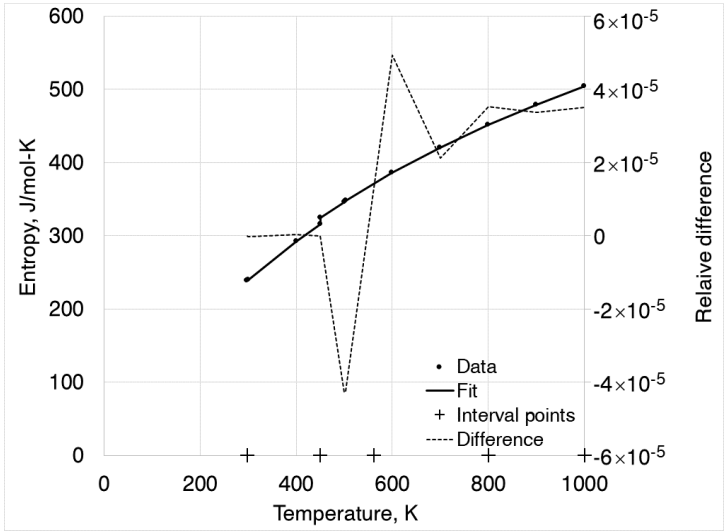


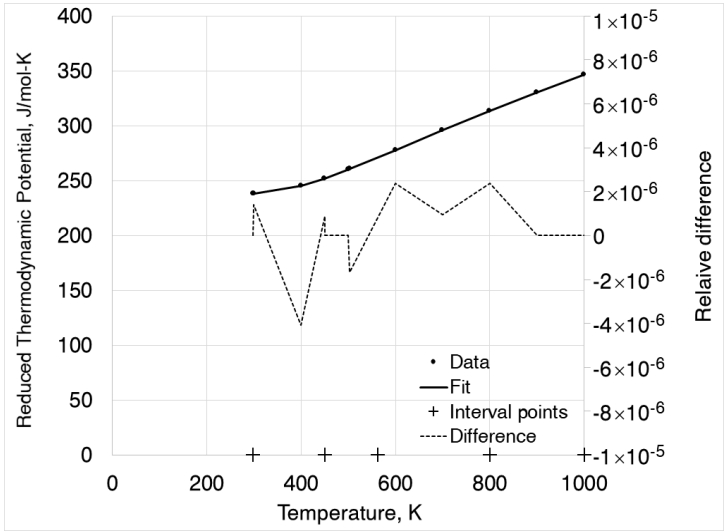
Tungsten Chloride, Alpha (WCl_6) – Cl6W1(cr)_01

MW = 396.57 g/mol – $\Delta_f H^o$(298.15 K) = -593.71 kJ/mol – S^o(298.15 K) = 238.488 J/mol-K – p^o = 0.1 MPa



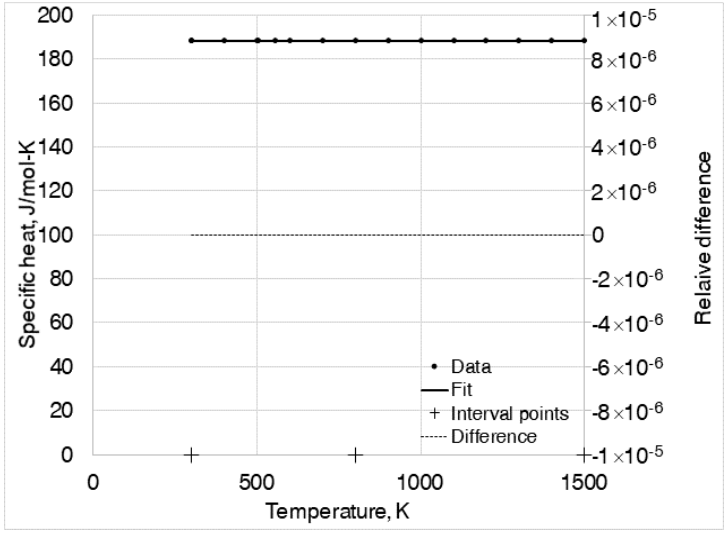


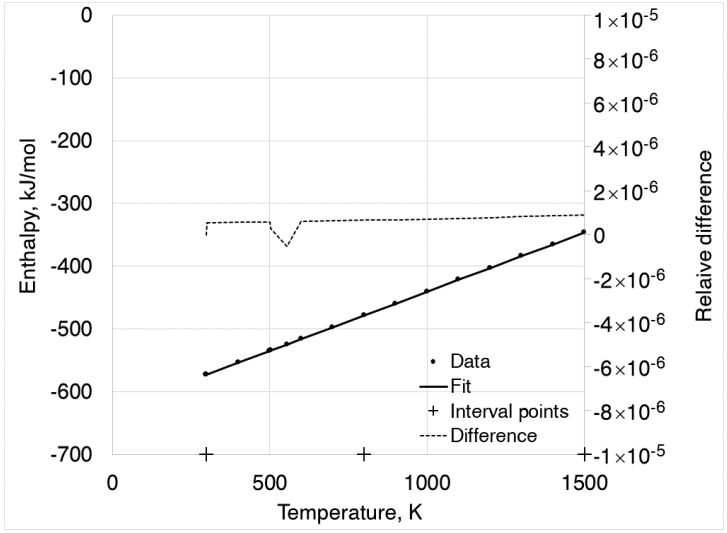


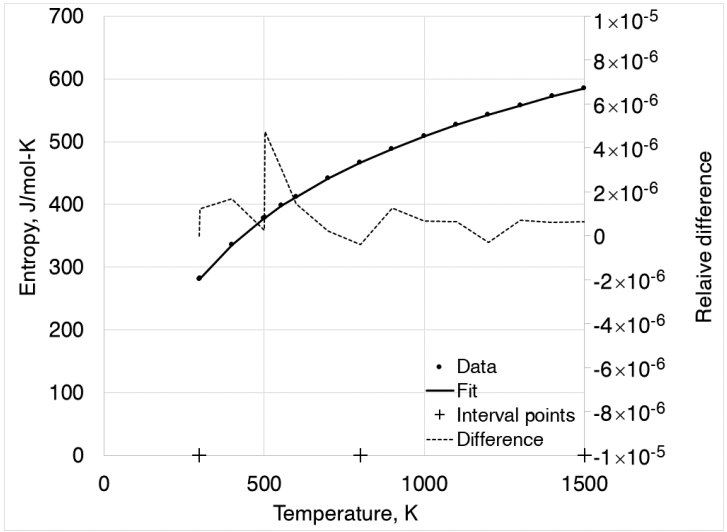


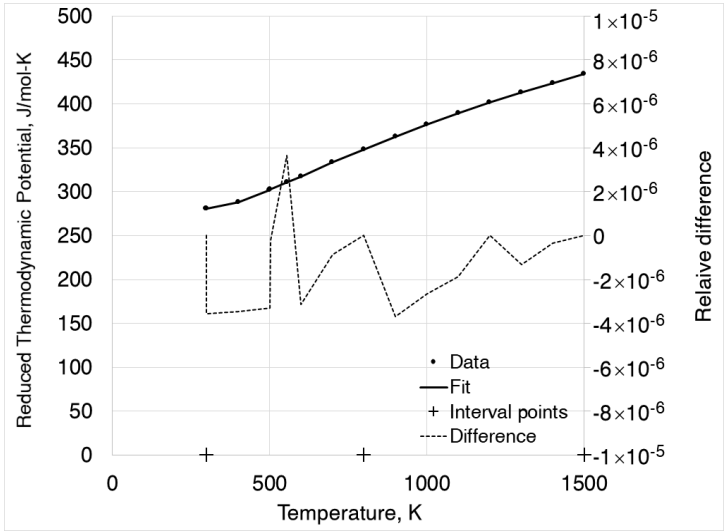
Tungsten Chloride, Beta (WCl_6) – Cl6W1(cr)_02

MW = 396.57 g/mol – $\Delta_f H^o$(298.15 K) = -572.877 kJ/mol – S^o(298.15 K) = 280.604 J/mol-K – p^o = 0.1 MPa



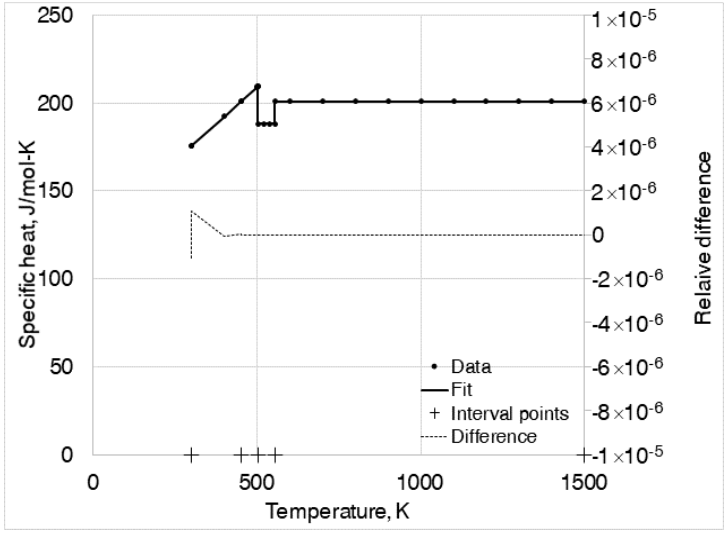


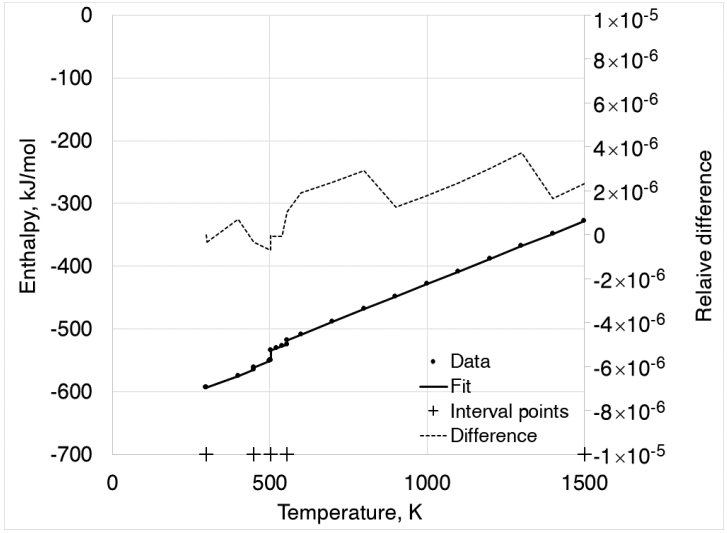


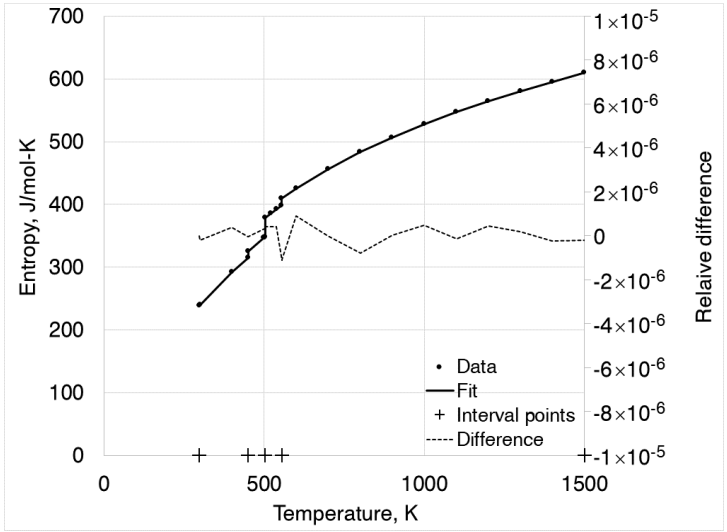


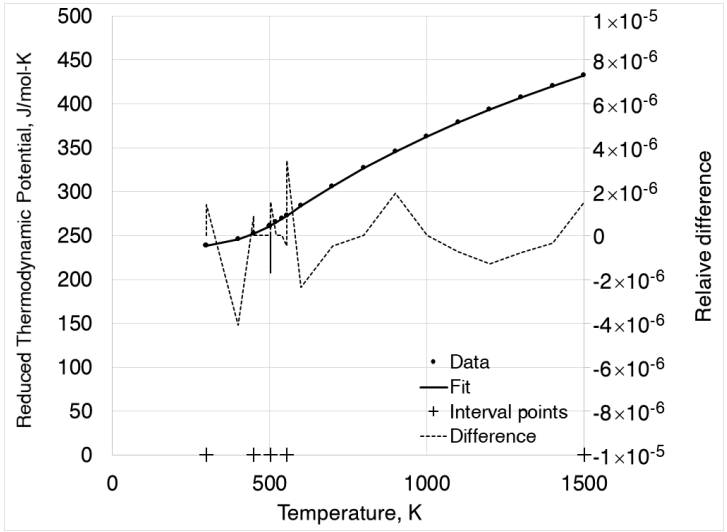
Tungsten Chloride (WCl_6) – Cl6W1(cr,l)

MW = 396.57 g/mol – $\Delta_f H^o$(298.15 K) = -593.71 kJ/mol – S^o(298.15 K) = 238.488 J/mol-K – p^o = 0.1 MPa



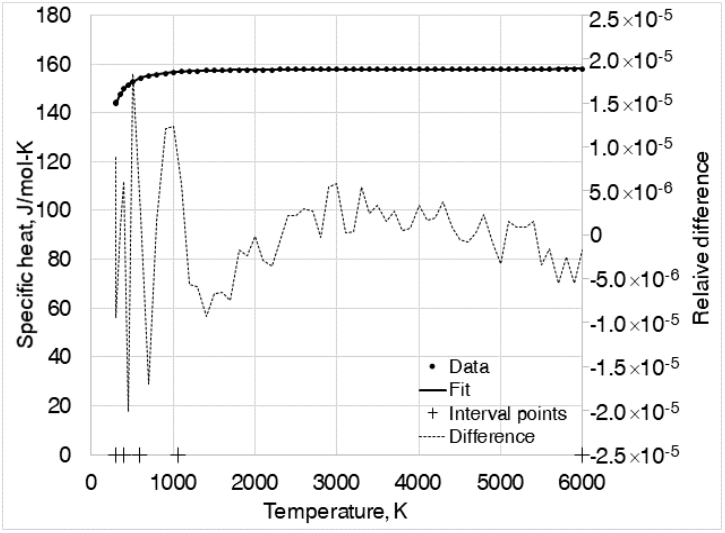


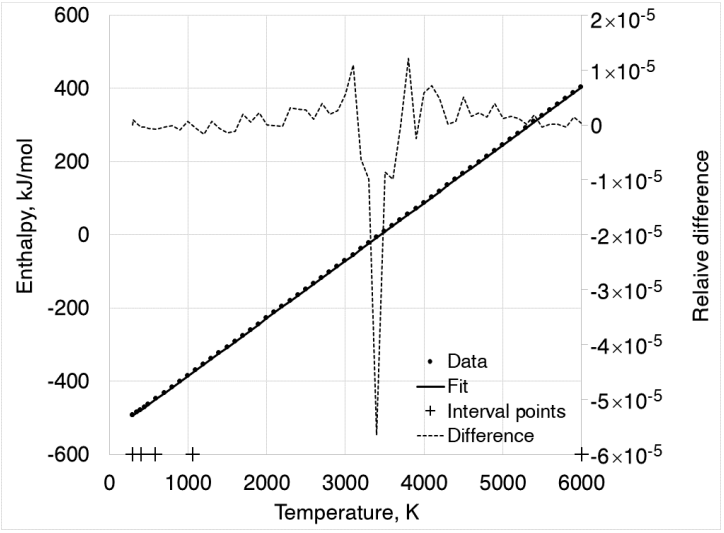


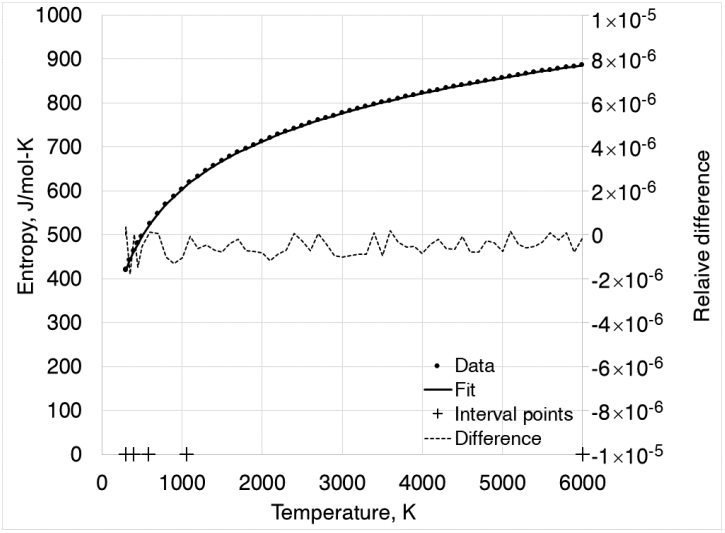


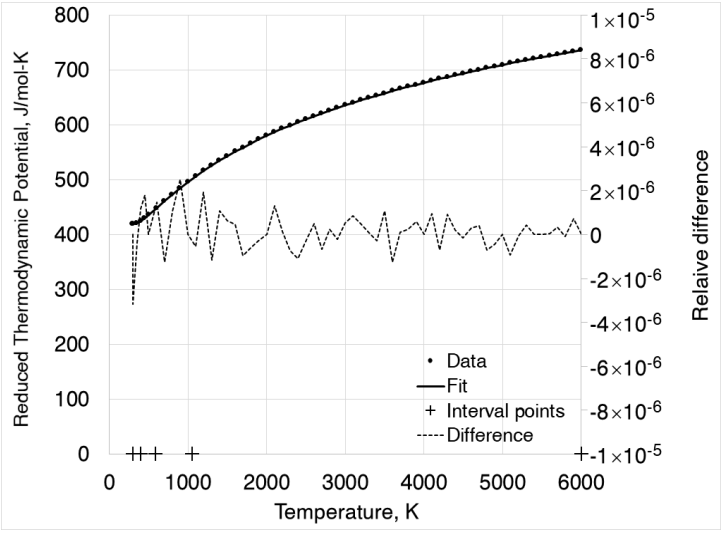
Tungsten Chloride (WCl_6) – Cl6W1(g)

MW = 396.57 g/mol – $\Delta_f H^o$(298.15 K) = -493.712 kJ/mol – S^o(298.15 K) = 419.166 J/mol-K – p^o = 0.1 MPa



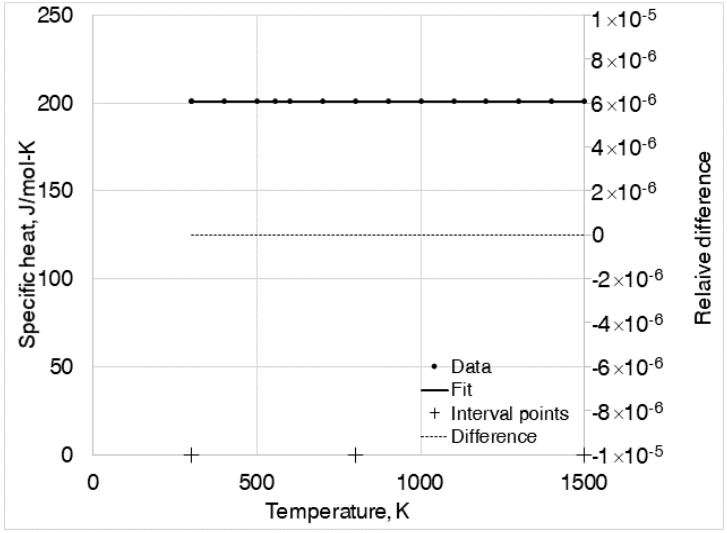


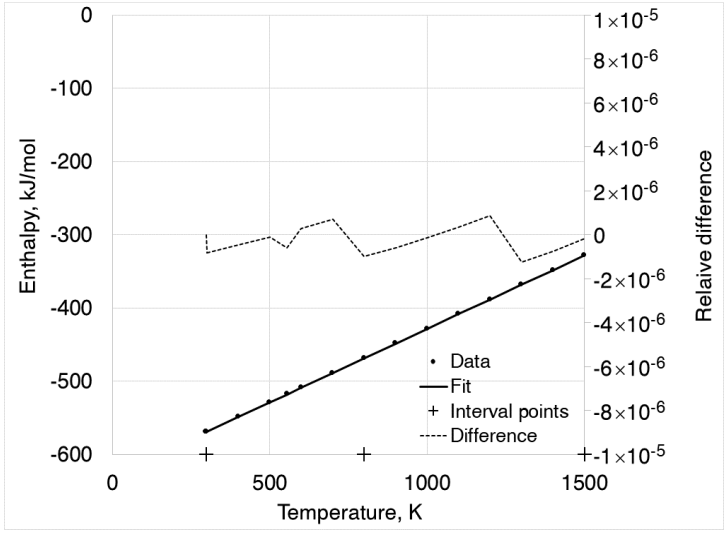


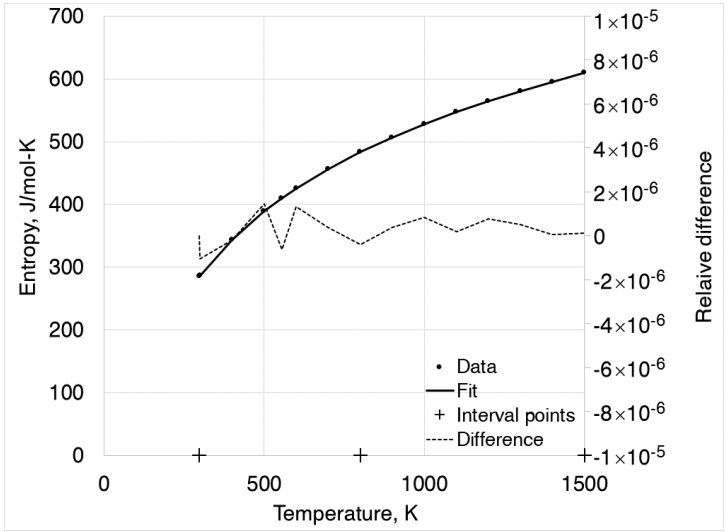


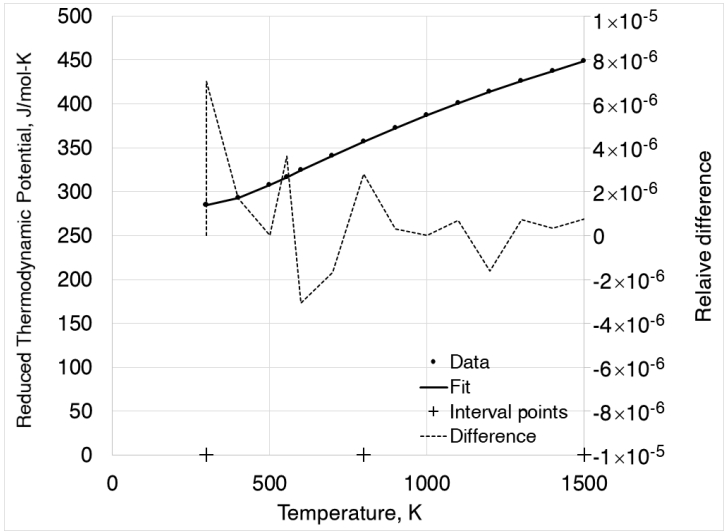
Tungsten Chloride (WCl_6) – Cl6W1(l)

MW = 396.57 g/mol – $\Delta_f H^o$(298.15 K) = -569.405 kJ/mol – S^o(298.15 K) = 284.868 J/mol-K – p^o = 0.1 MPa



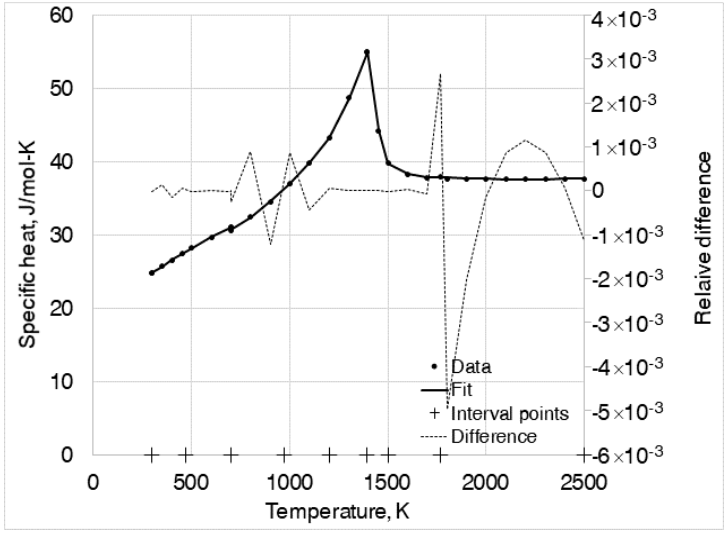


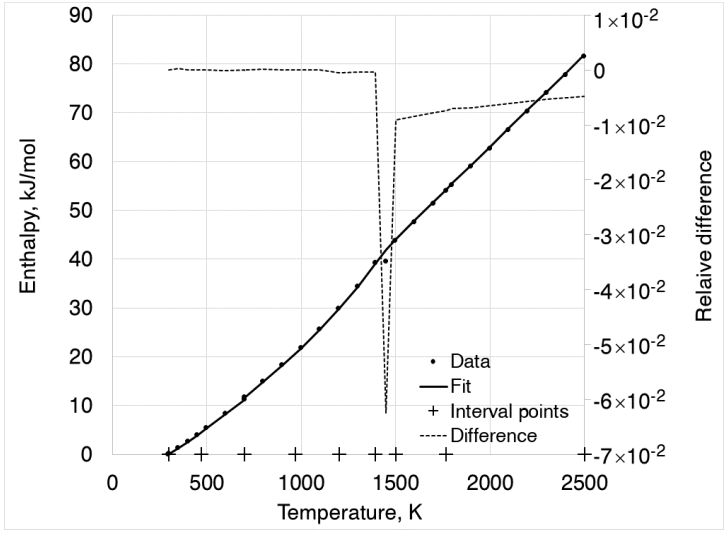


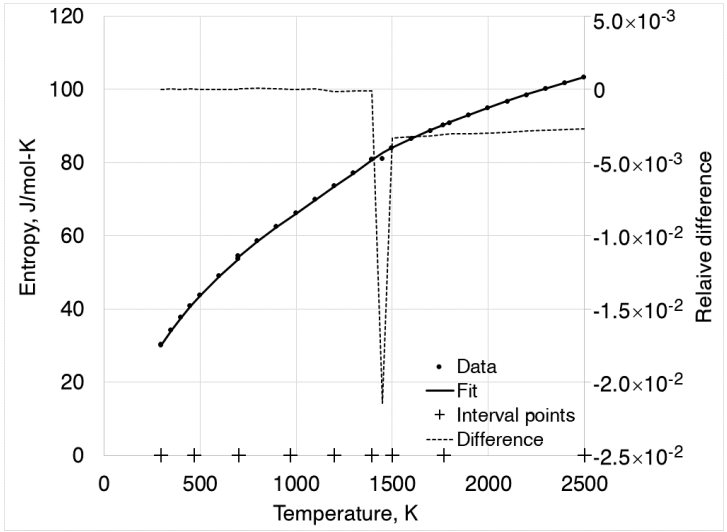


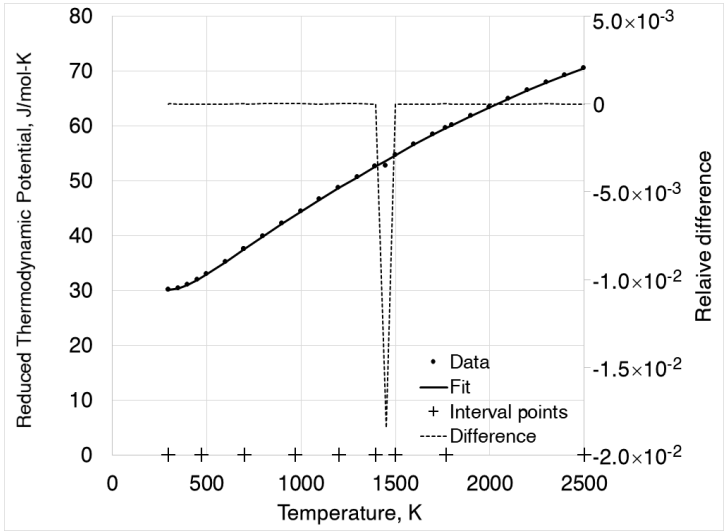
Cobalt (Co) – Co1(cr)

MW = 58.9312 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 30.067 J/mol-K – p^o = 0.1 MPa



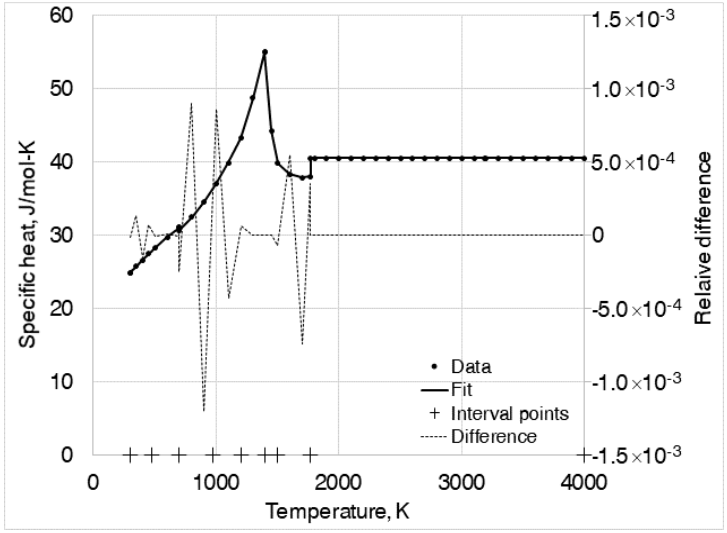


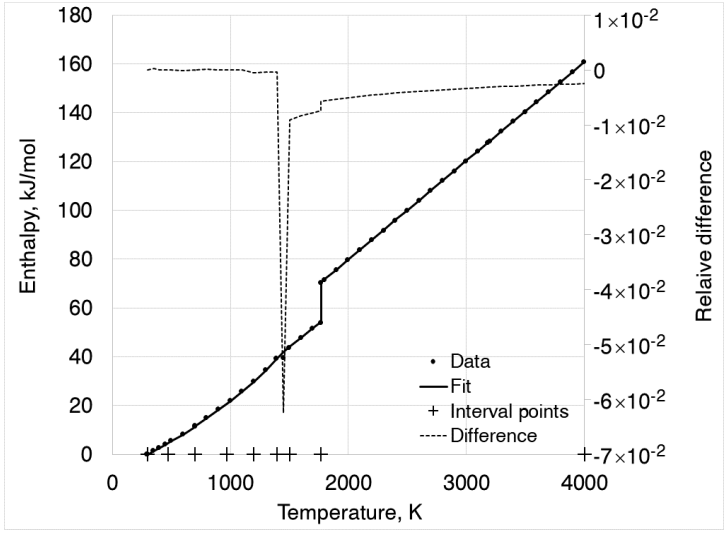


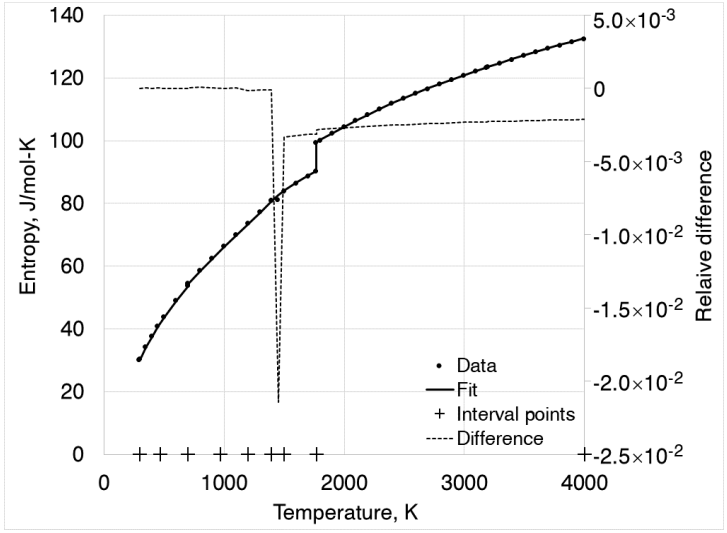


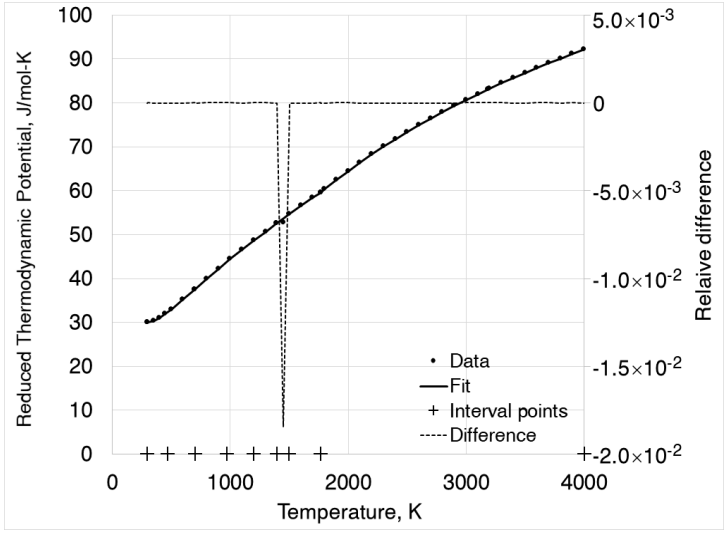
Cobalt (Co) – Co1(cr,l)

MW = 58.9312 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 30.067 J/mol-K – p^o = 0.1 MPa



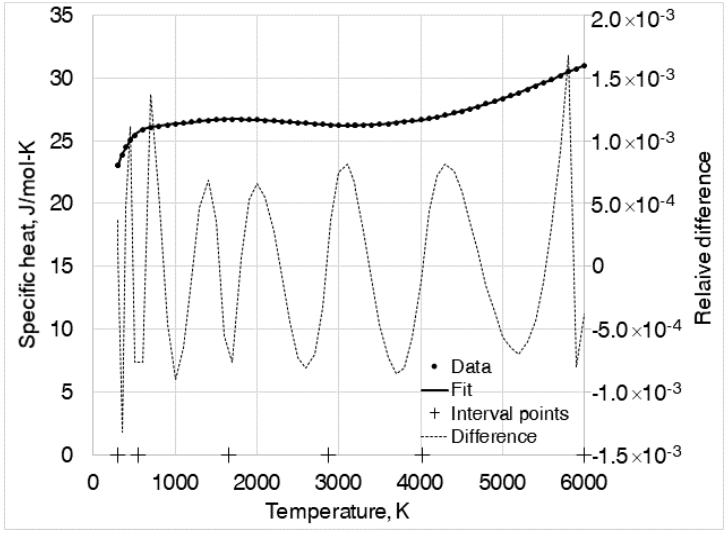


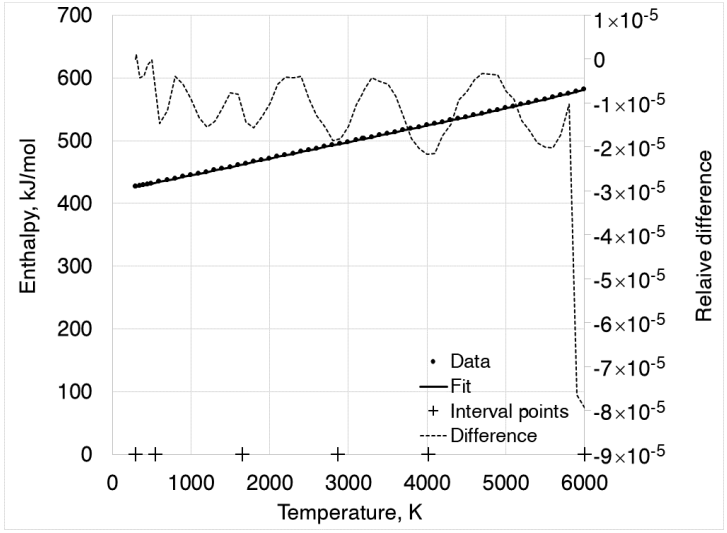


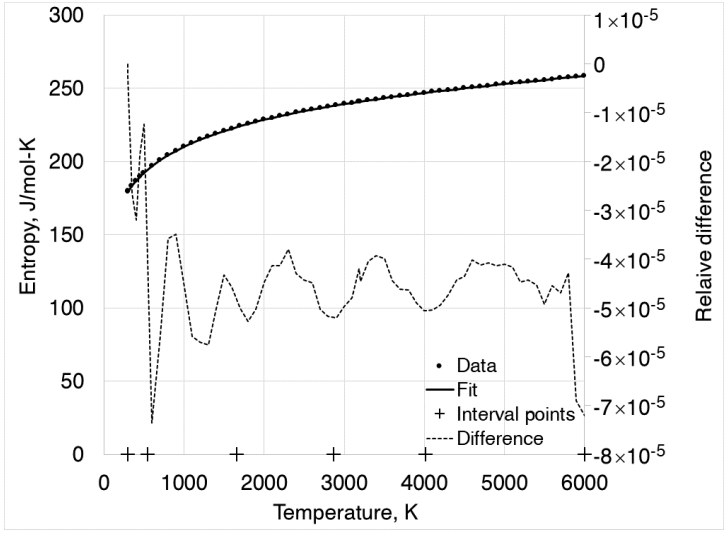


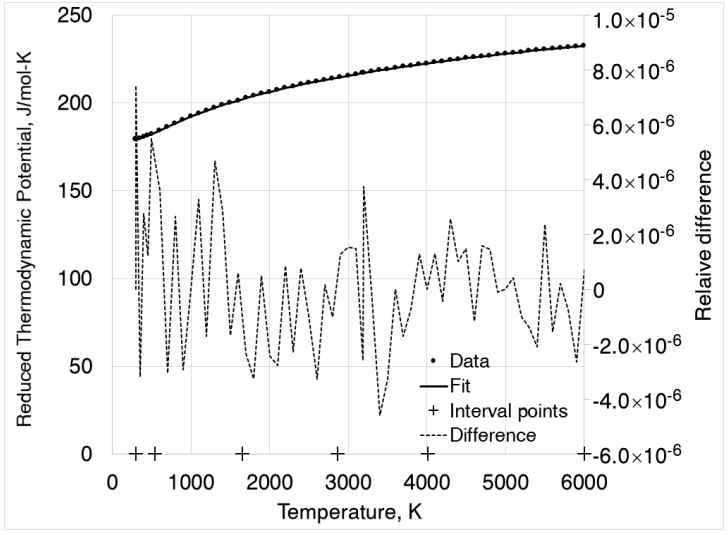
Cobalt (Co) – Co1(g)

MW = 58.9312 g/mol – $\Delta_f H^o$(298.15 K) = 426.676 kJ/mol – S^o(298.15 K) = 179.518 J/mol-K – p^o = 0.1 MPa



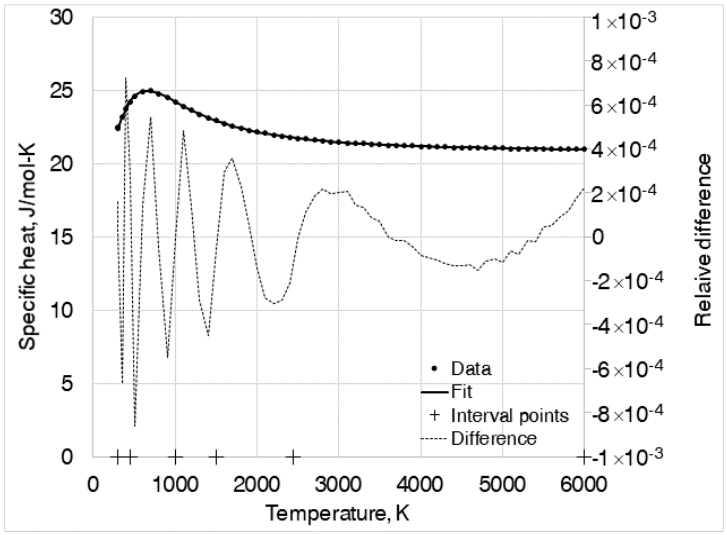


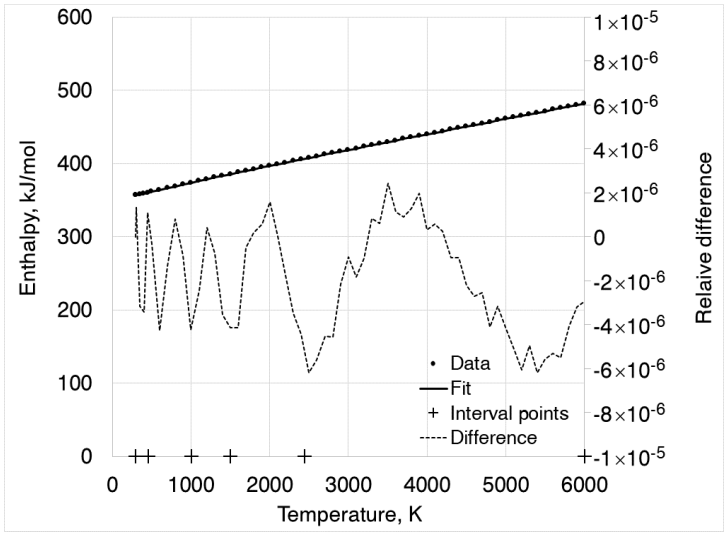


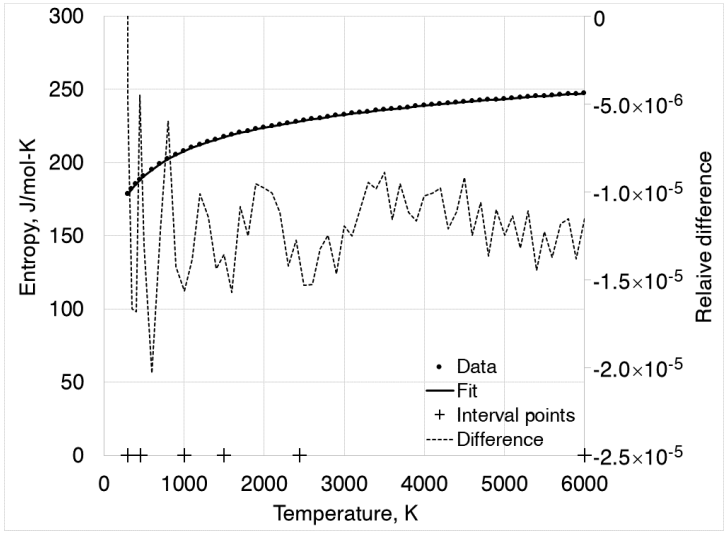


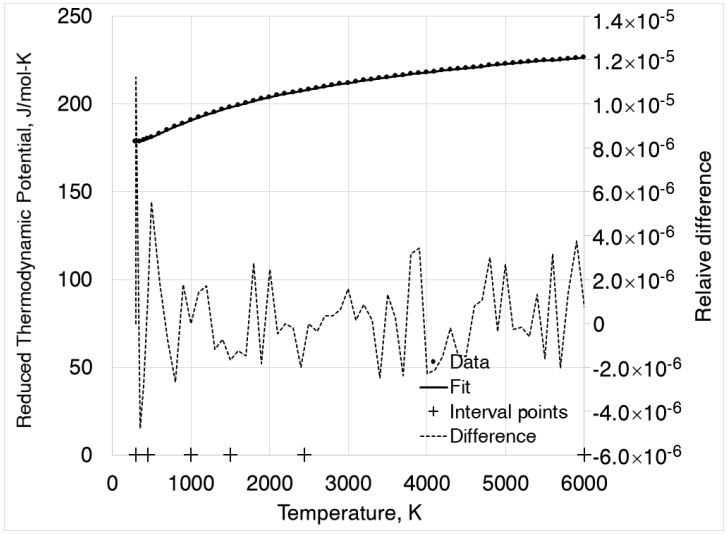
Cobalt, Ion (Co^-) – Co1-(g)

MW = 58.9317 g/mol – $\Delta_f H^o$(298.15 K) = 356.649 kJ/mol – S^o(298.15 K) = 178.412 J/mol-K – p^o = 0.1 MPa



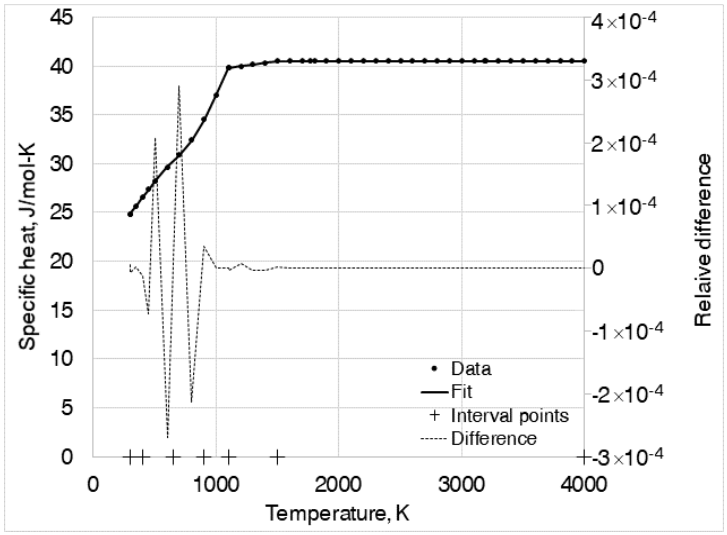


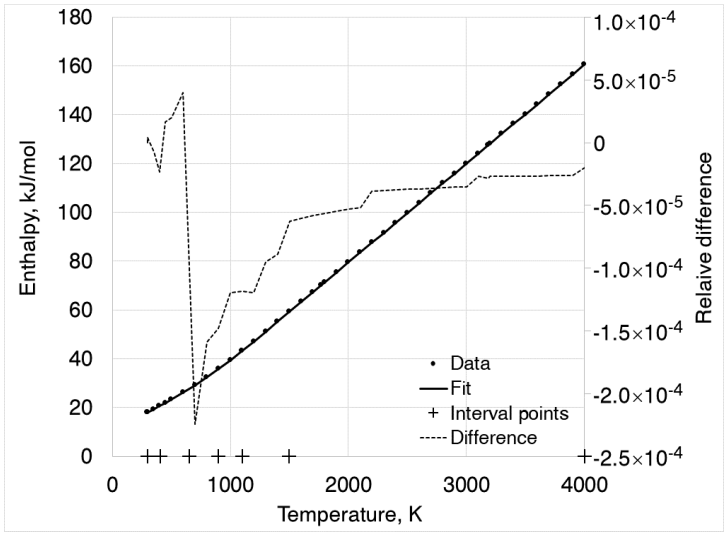


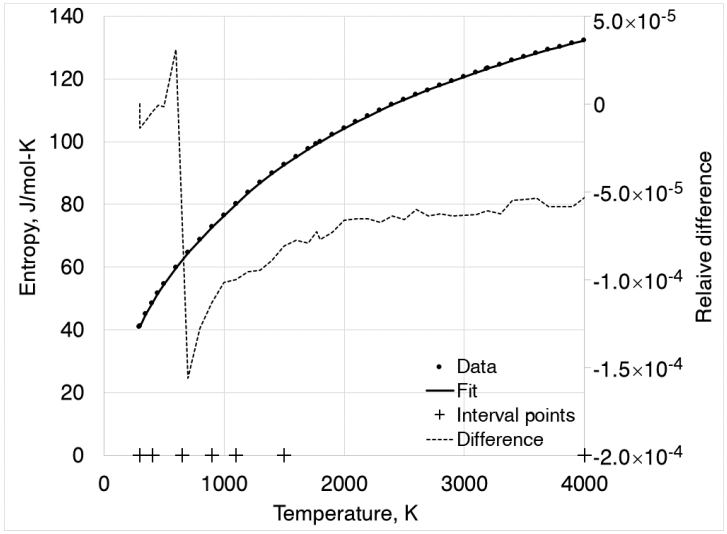


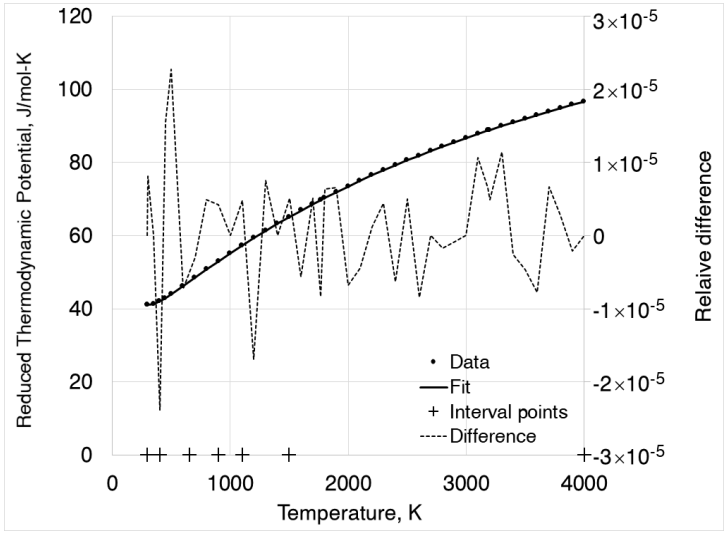
Cobalt (Co) – Co1(l)

MW = 58.9312 g/mol – $\Delta_f H^o$(298.15 K) = 18.004 kJ/mol – S^o(298.15 K) = 41 J/mol-K – p^o = 0.1 MPa



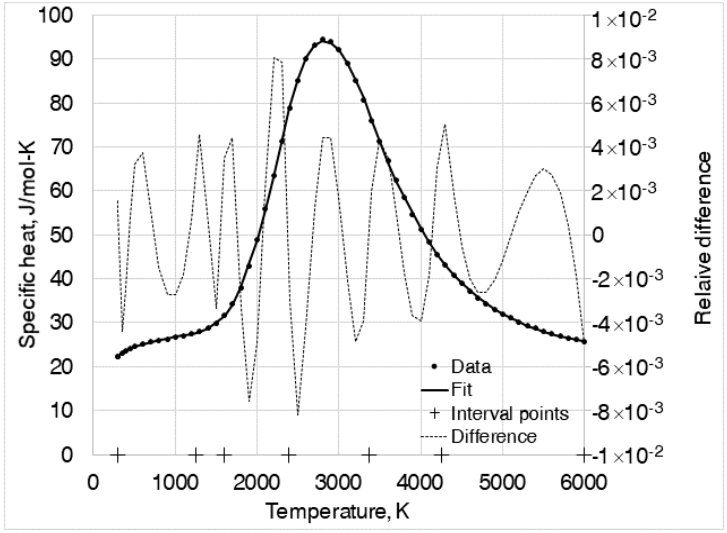


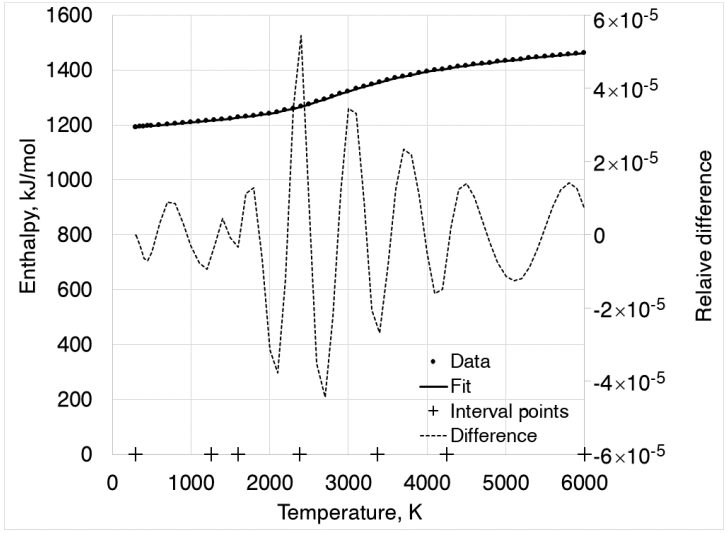


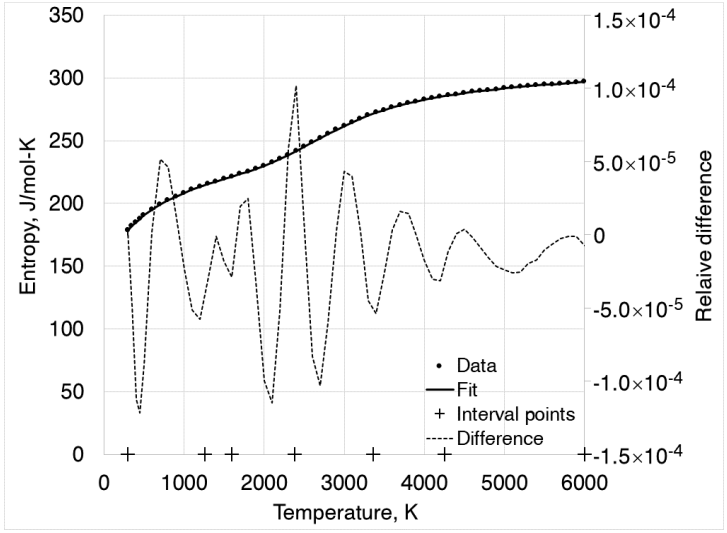


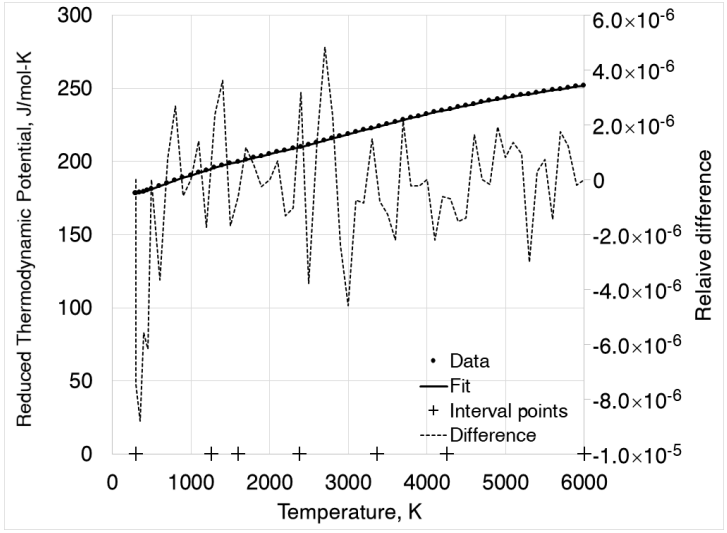
Cobalt, Ion (Co^+) – Co1+(g)

MW = 58.9307 g/mol – $\Delta_f H^o$(298.15 K) = 1191.597 kJ/mol – S^o(298.15 K) = 178.346 J/mol-K – p^o = 0.1 MPa



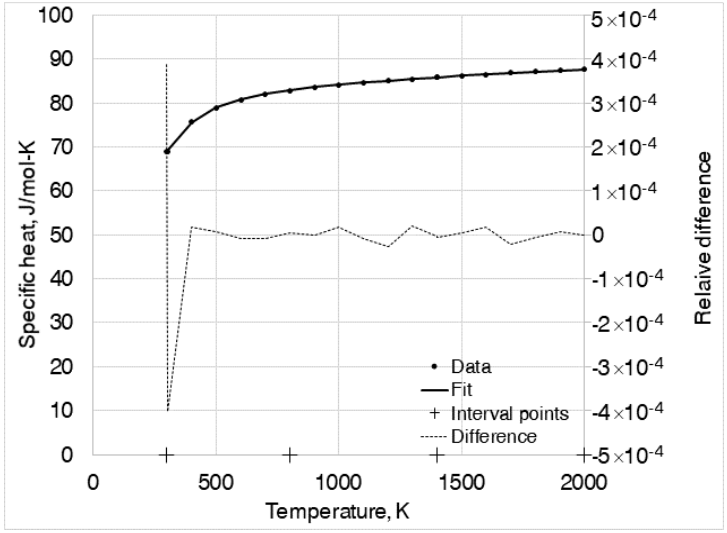


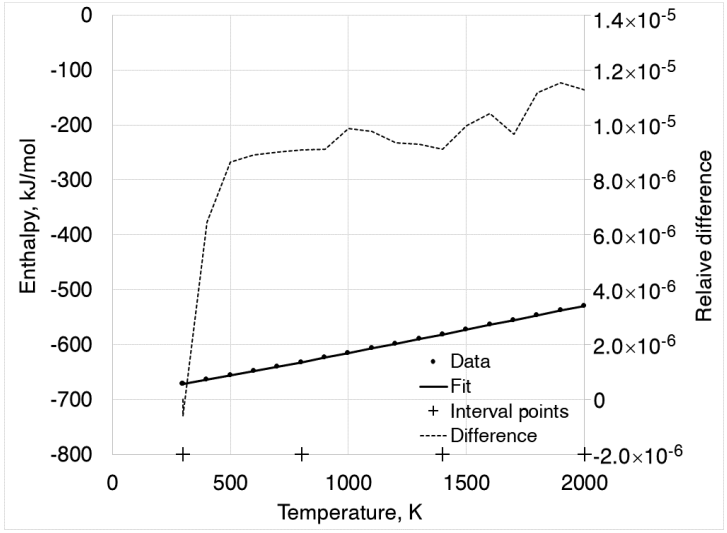


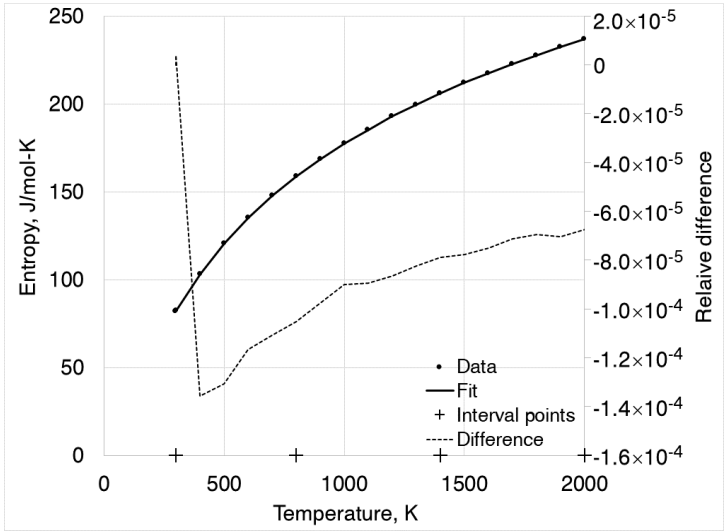


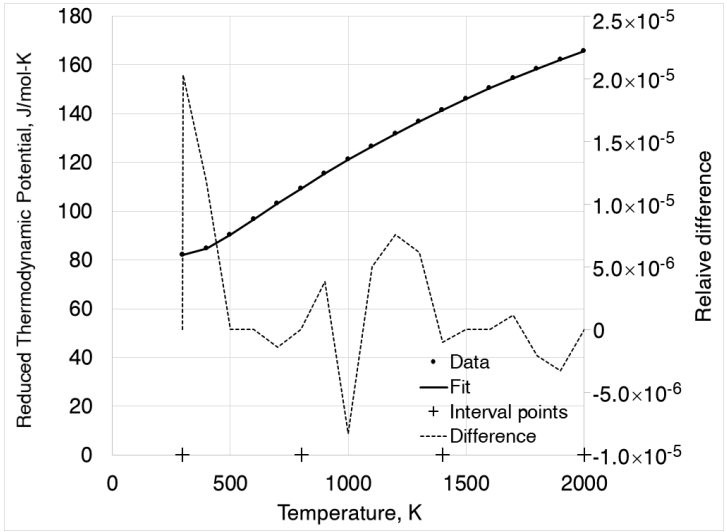
Cobalt Fluoride (CoF_2) – Co1F2(cr)

MW = 96.928 g/mol – $\Delta_f H^o$(298.15 K) = -671.532 kJ/mol – S^o(298.15 K) = 82.026 J/mol-K – p^o = 0.1 MPa



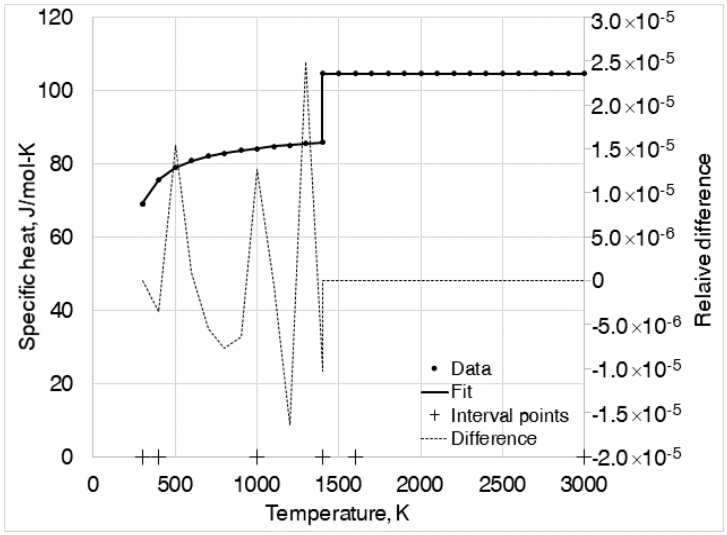


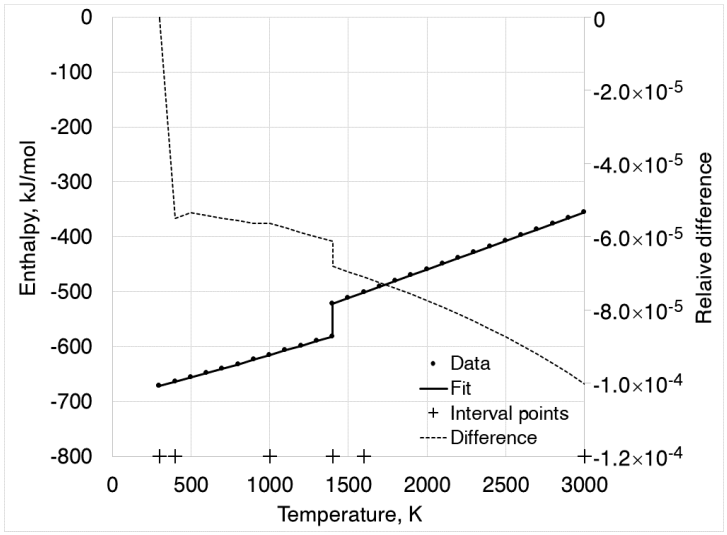


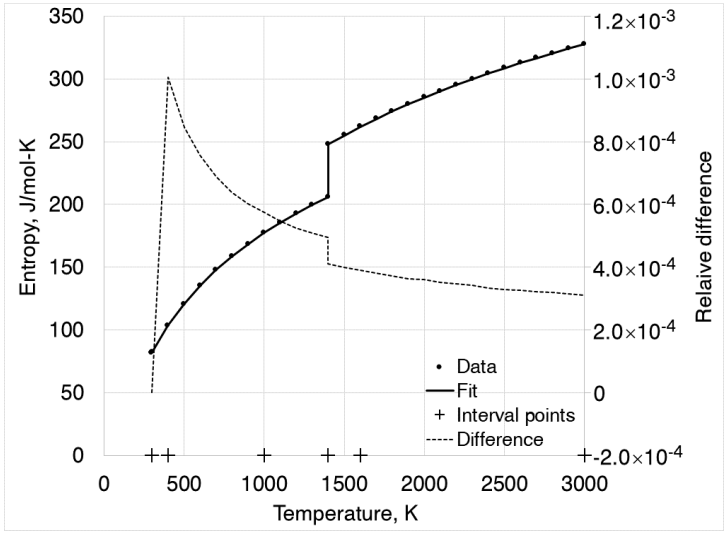


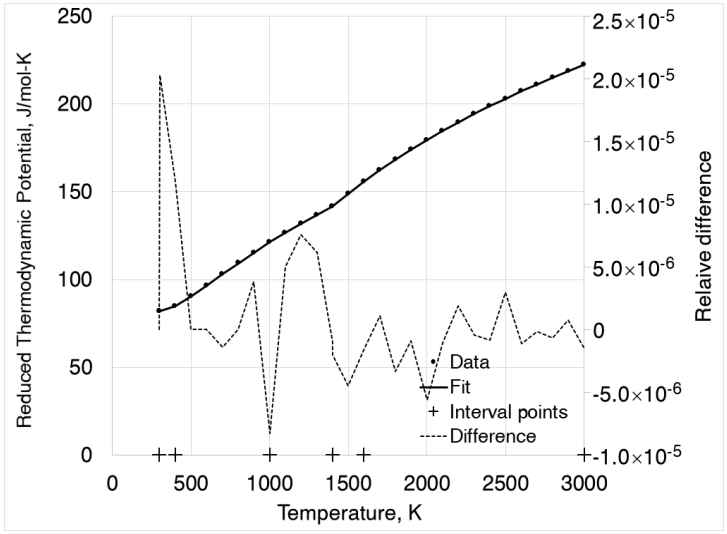
Cobalt Fluoride (CoF_2) – Co1F2(cr,l)

MW = 96.928 g/mol – $\Delta_f H^o$(298.15 K) = -671.532 kJ/mol – S^o(298.15 K) = 82.026 J/mol-K – p^o = 0.1 MPa



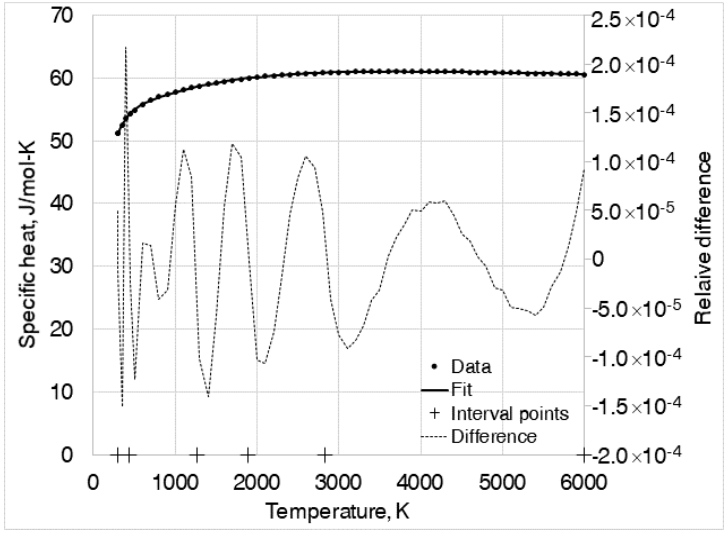


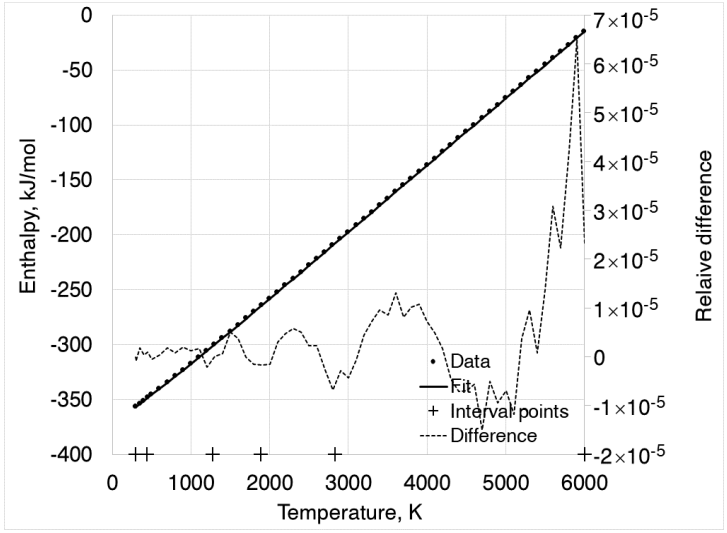


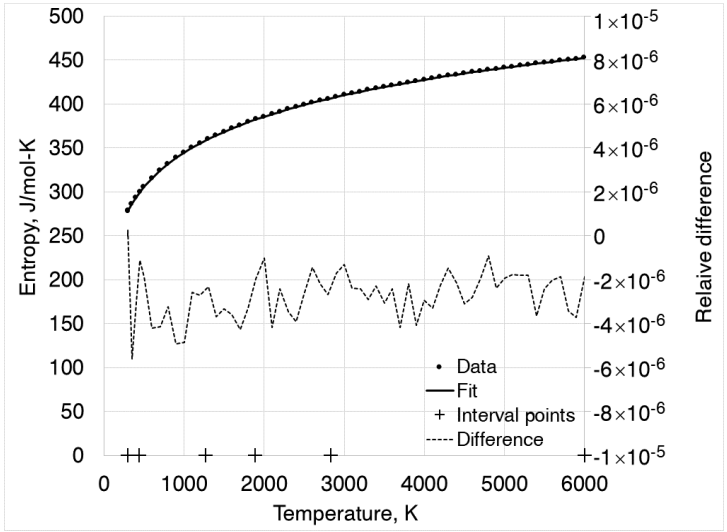


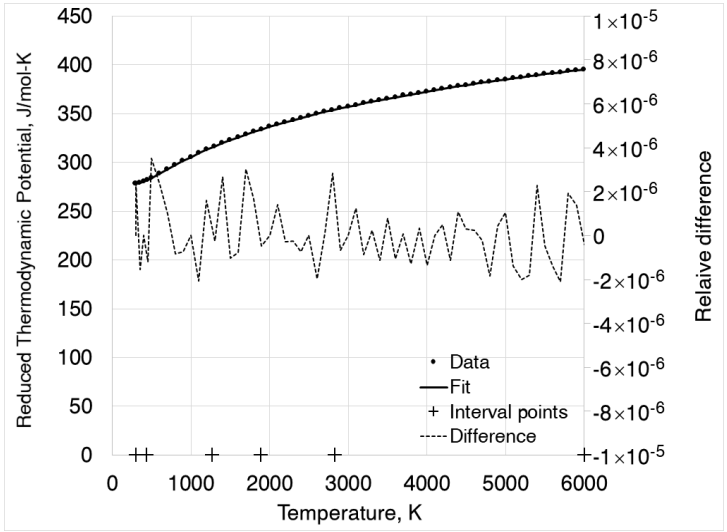
Cobalt Fluoride (CoF_2) – Co1F2(g)

MW = 96.928 g/mol – $\Delta_f H^o$(298.15 K) = -356.477 kJ/mol – S^o(298.15 K) = 278.042 J/mol-K – p^o = 0.1 MPa



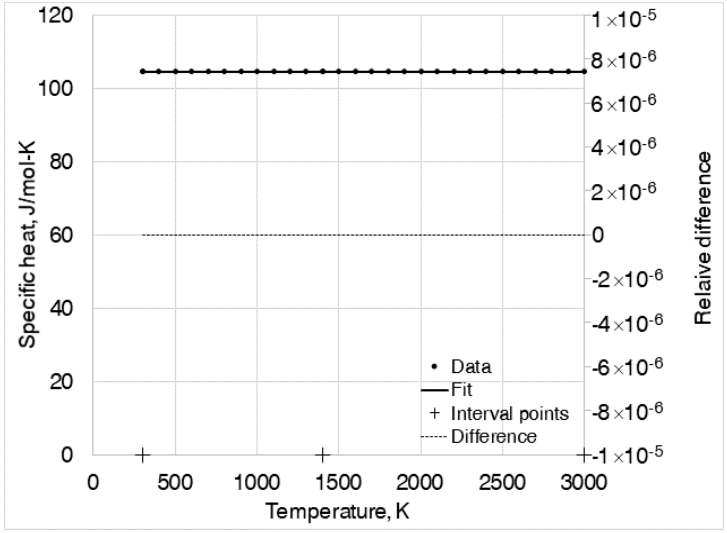


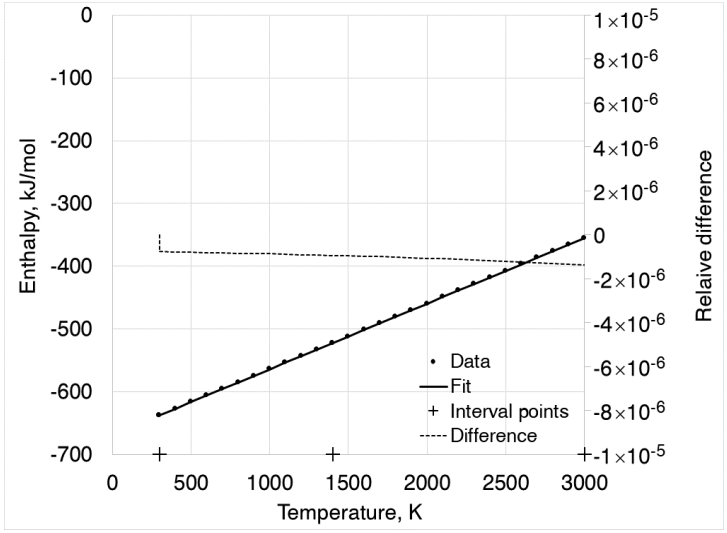


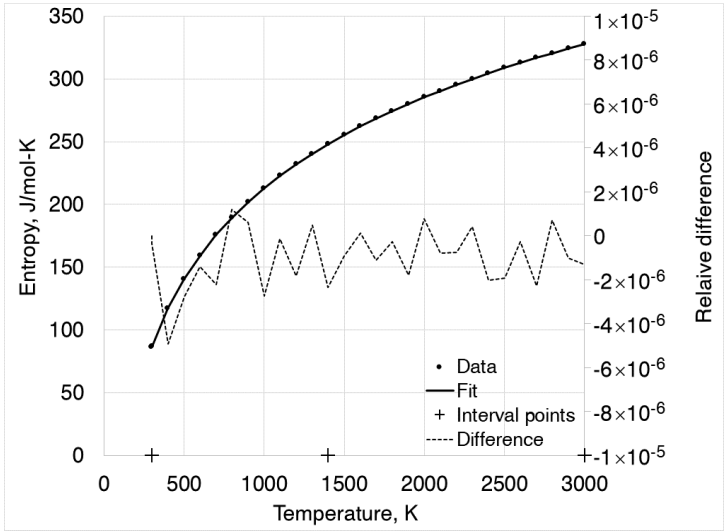


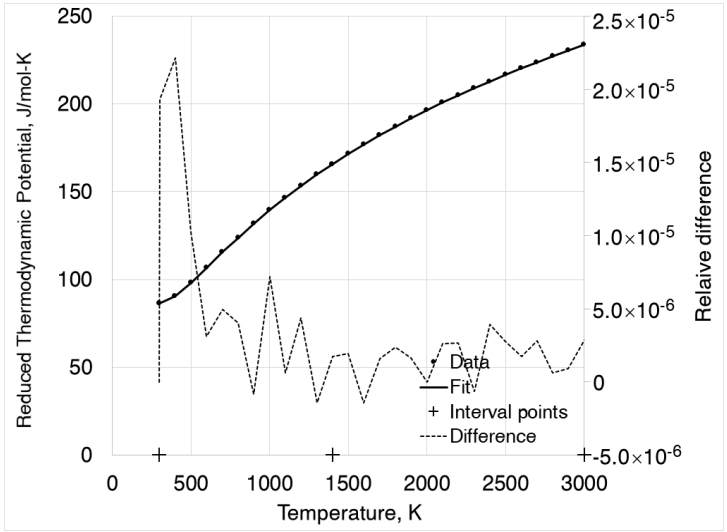
Cobalt Fluoride (CoF_2) – Co1F2(l)

MW = 96.928 g/mol – $\Delta_f H^o$(298.15 K) = -637.691 kJ/mol – S^o(298.15 K) = 86.238 J/mol-K – p^o = 0.1 MPa



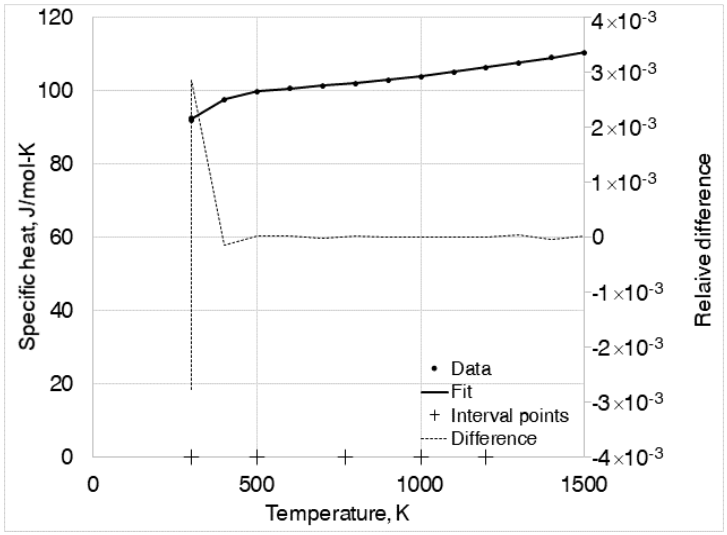


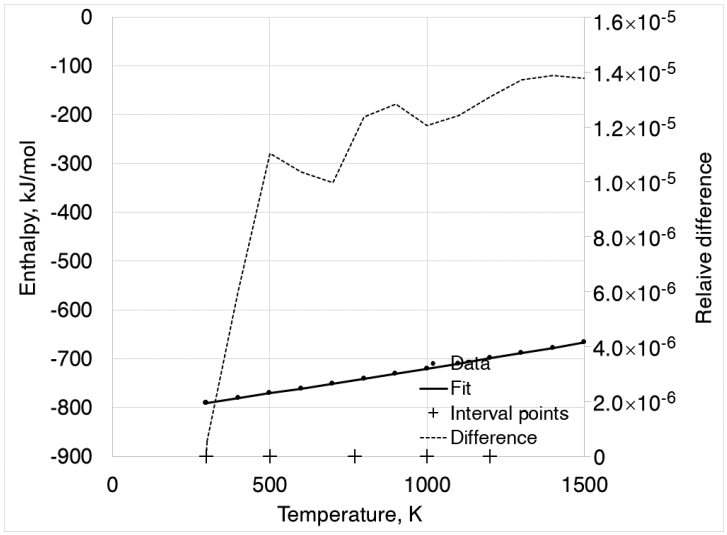


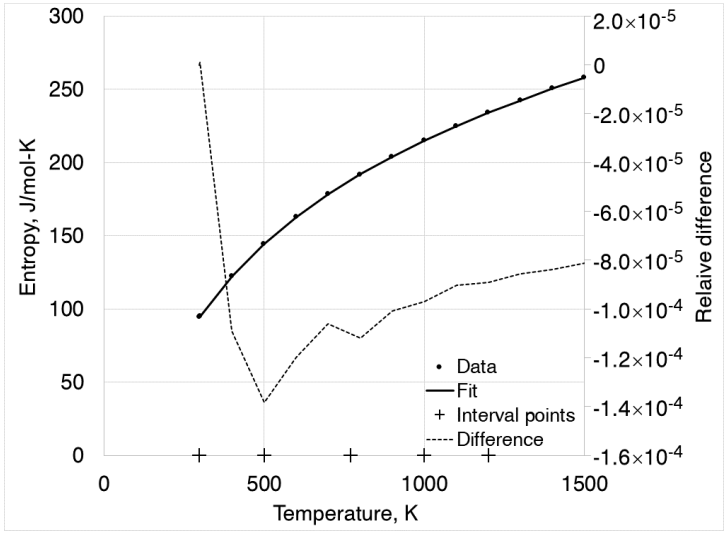


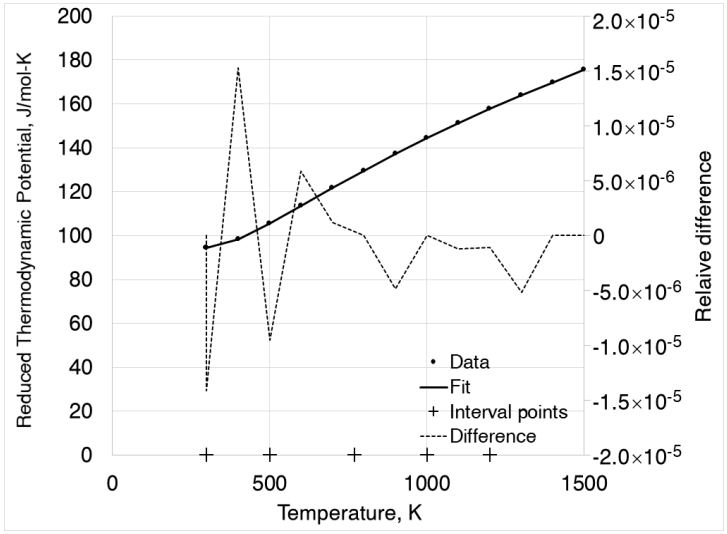
Cobalt Fluoride (CoF_3) – Co1F3(cr)

MW = 115.9264 g/mol – $\Delta_f H^o$(298.15 K) = -790.358 kJ/mol – S^o(298.15 K) = 94.558 J/mol-K – p^o = 0.1 MPa



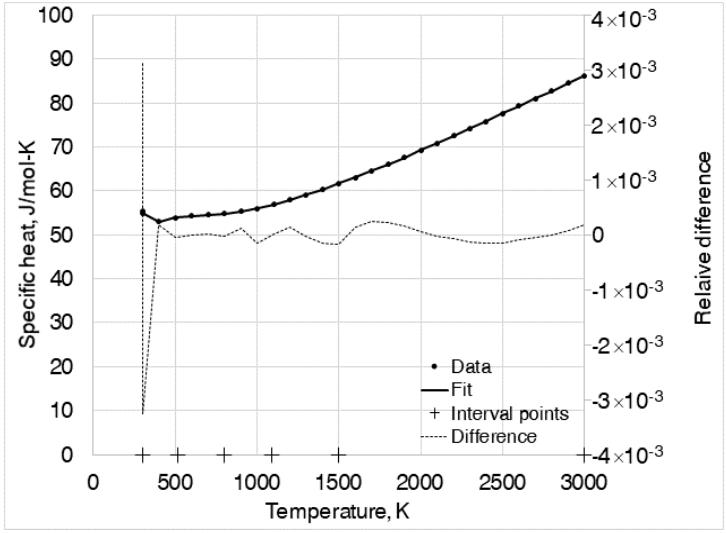


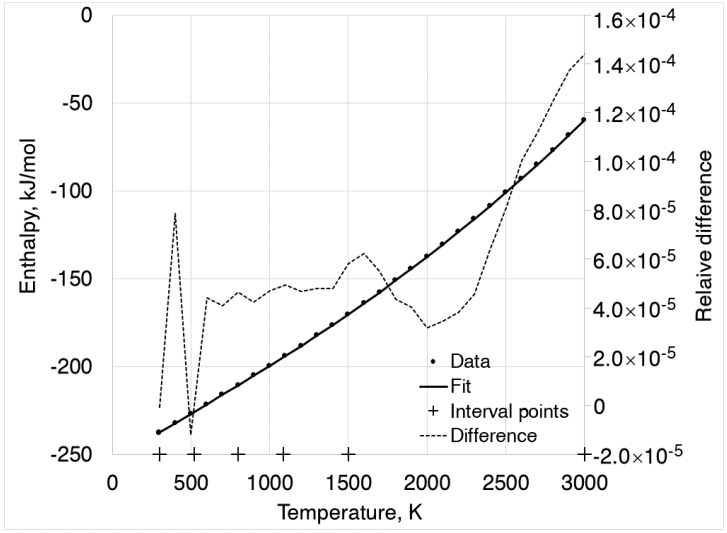


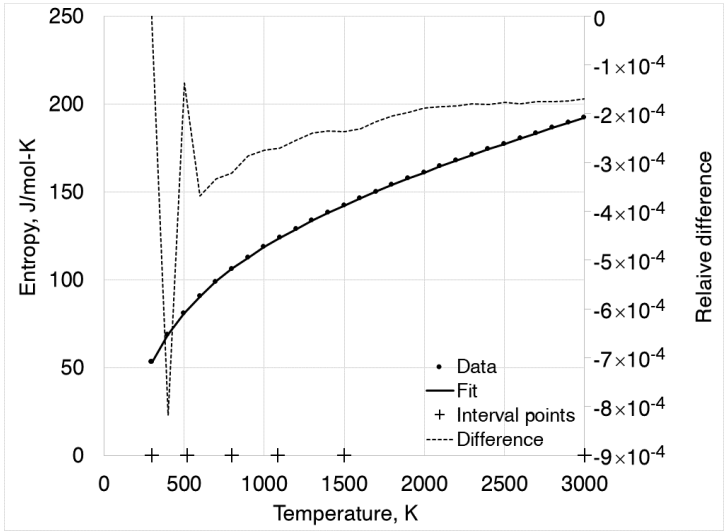


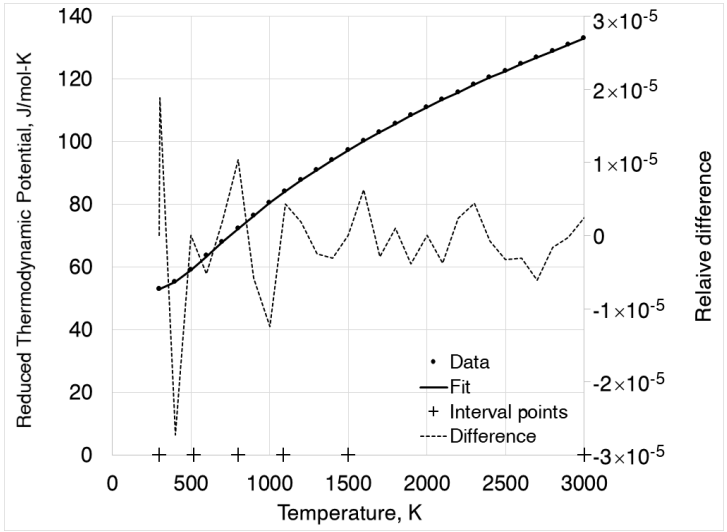
Cobalt Oxide (CoO) – Co1O1(cr)

MW = 74.9306 g/mol – $\Delta_f H^o$(298.15 K) = -237.735 kJ/mol – S^o(298.15 K) = 52.993 J/mol-K – p^o = 0.1 MPa



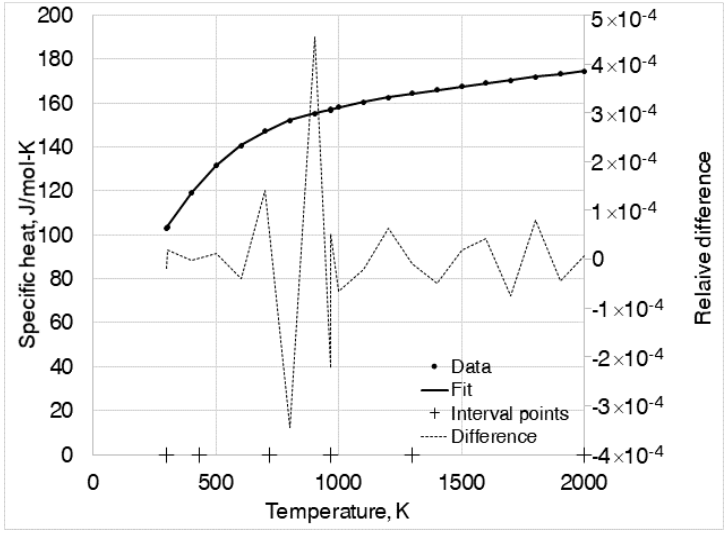


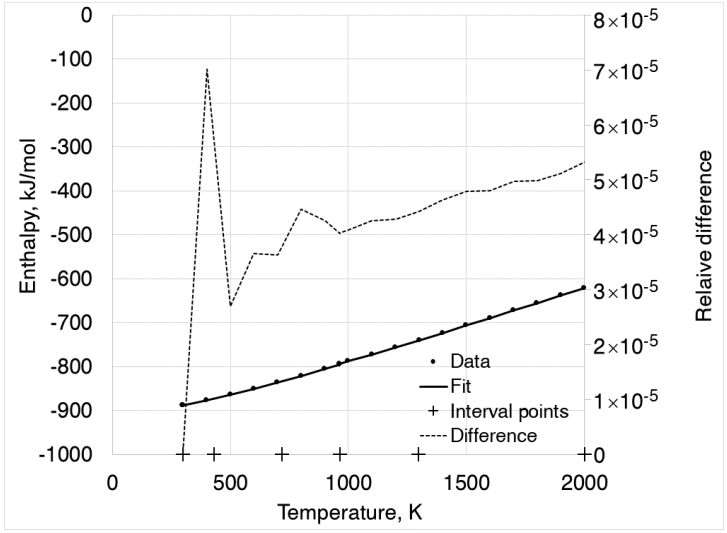


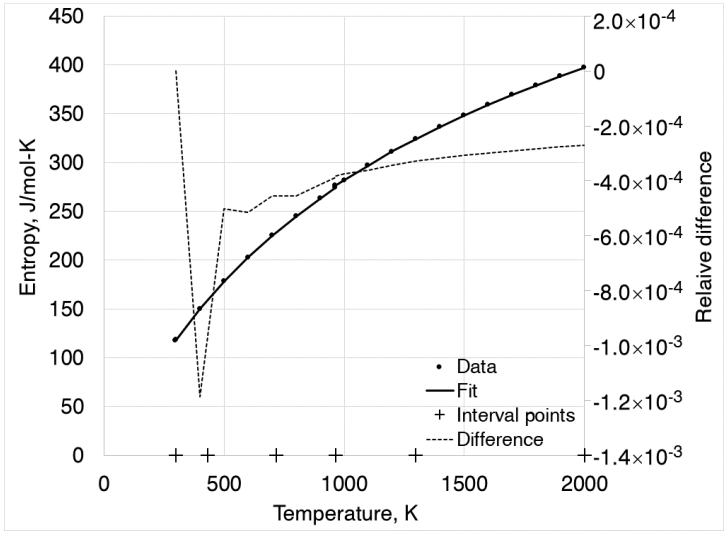


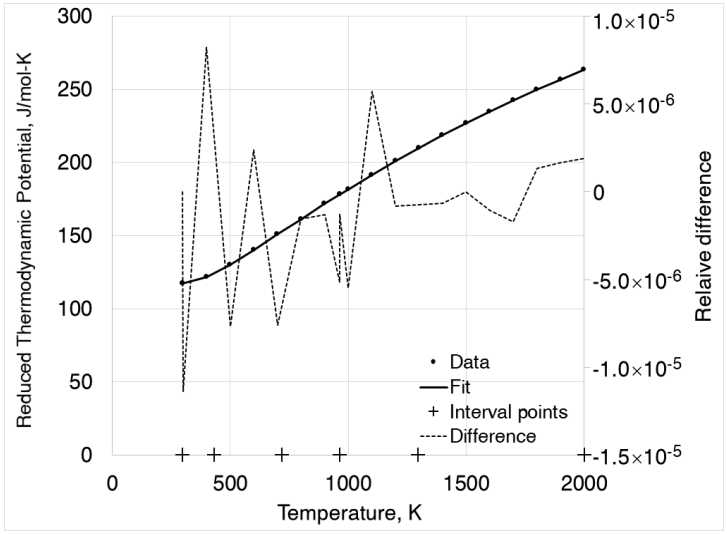
Cobalt Sulfate ($CoSO_4$) – Co1O4S1(cr)

MW = 154.99 g/mol – $\Delta_f H^o$(298.15 K) = -888.263 kJ/mol – S^o(298.15 K) = 117.377 J/mol-K – p^o = 0.1 MPa



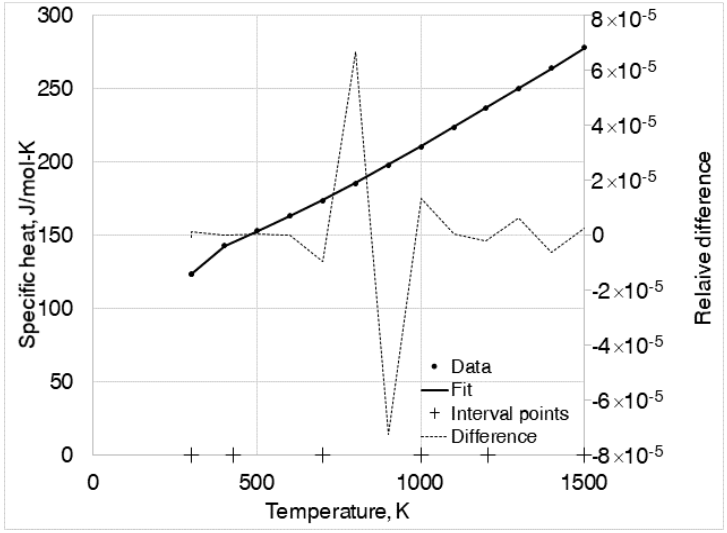


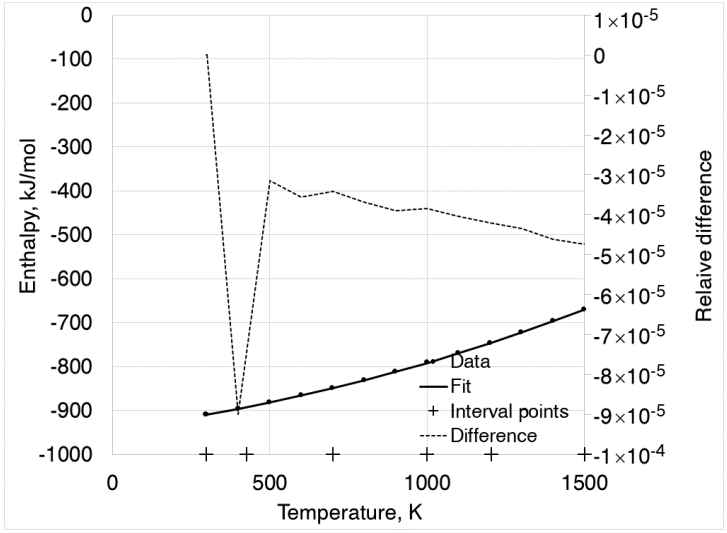


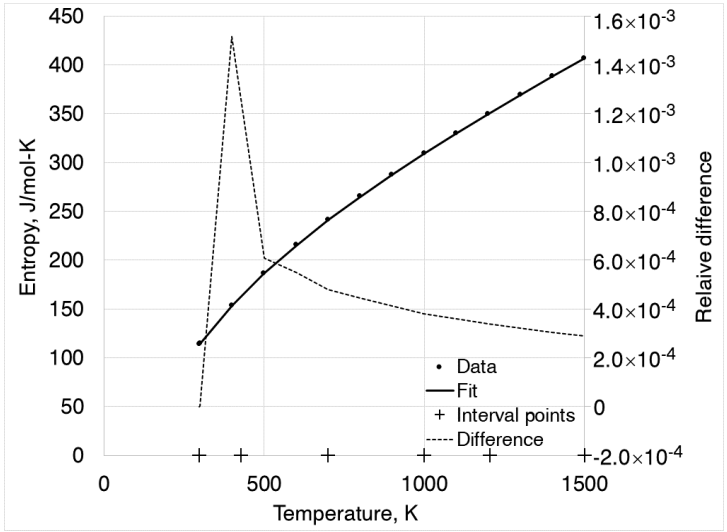


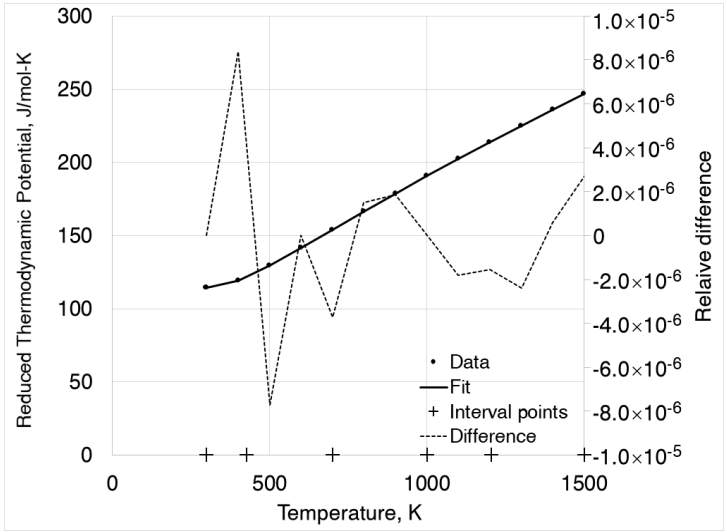
Cobalt Oxide (Co_3O_4) – Co3O4(cr)

MW = 240.7912 g/mol – $\Delta_f H^o$(298.15 K) = -910.02 kJ/mol – S^o(298.15 K) = 114.286 J/mol-K – p^o = 0.1 MPa



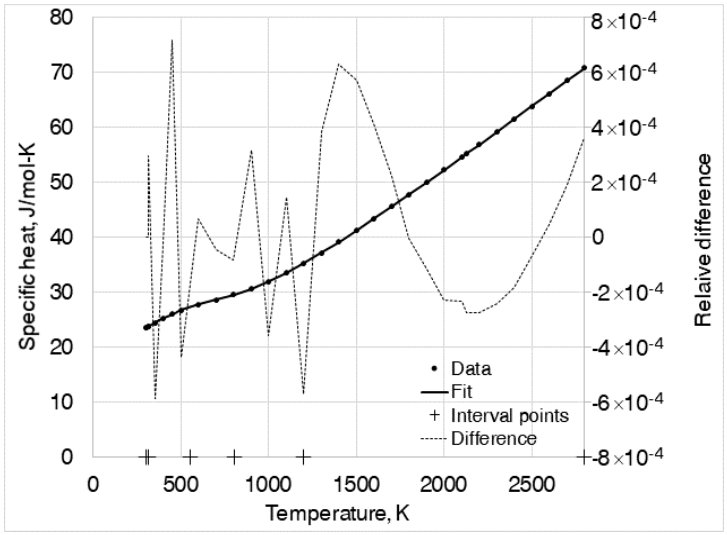


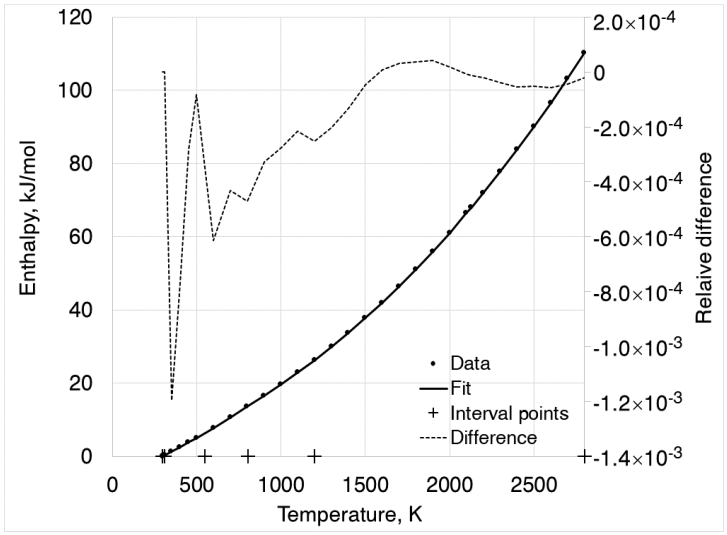


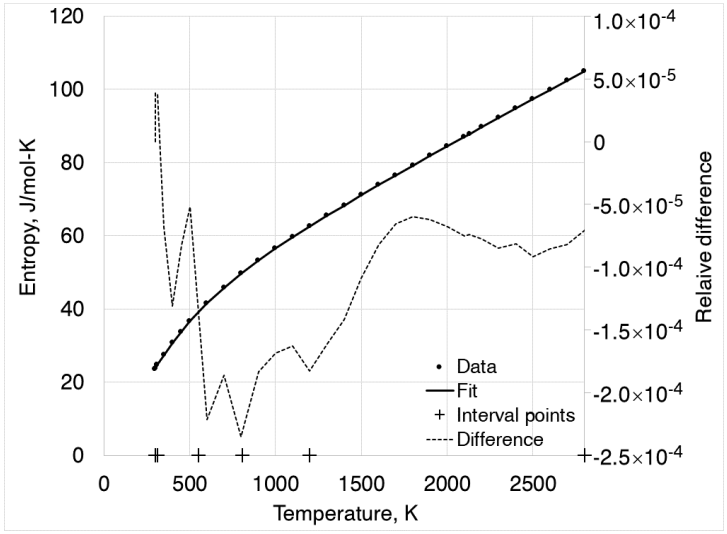


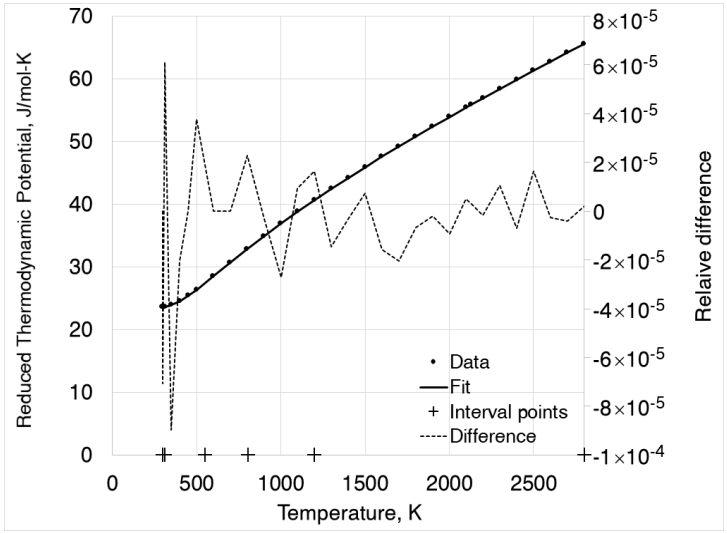
Chromium (Cr) – Cr1(cr)

MW = 51.996 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 23.618 J/mol-K – p^o = 0.1 MPa



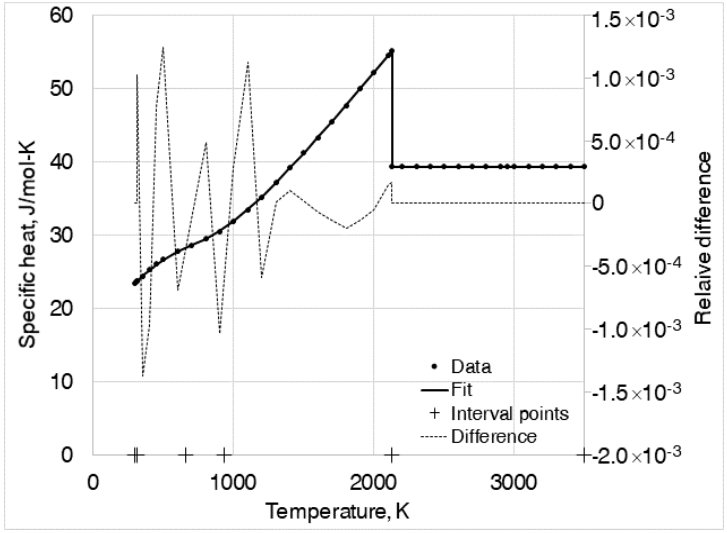


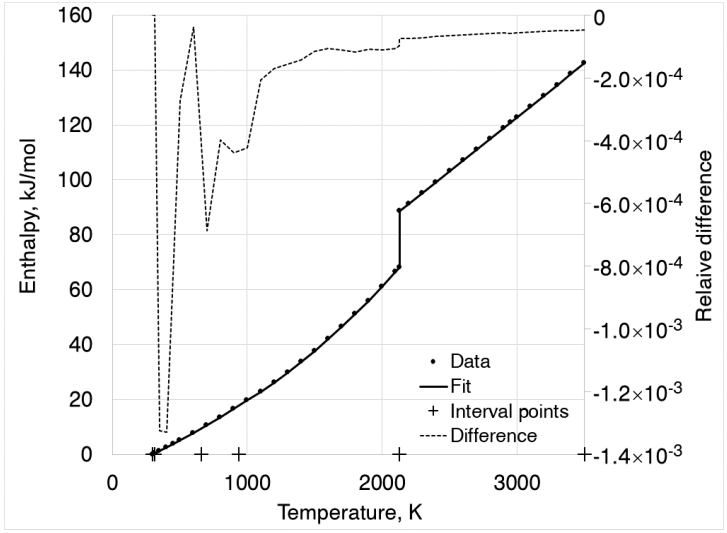


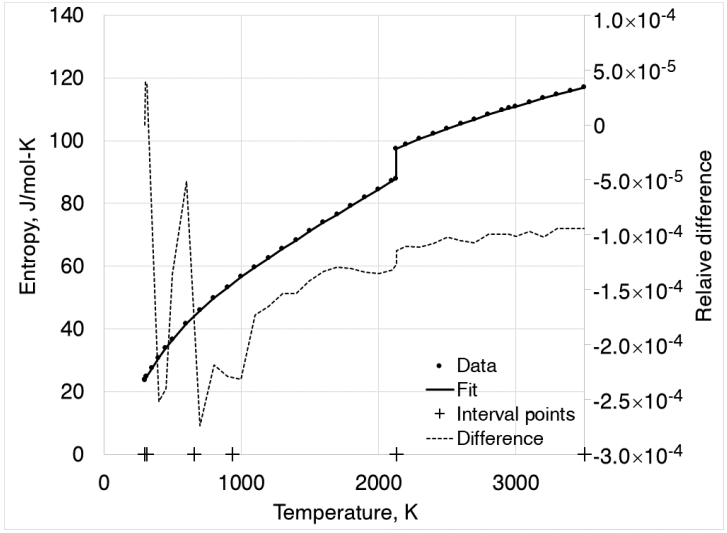


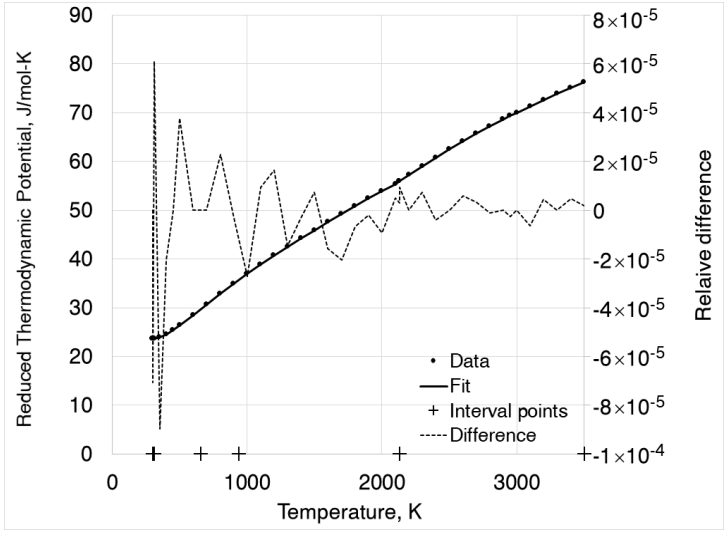
Chromium (Cr) – Cr1(cr,l)

MW = 51.996 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 23.618 J/mol-K – p^o = 0.1 MPa



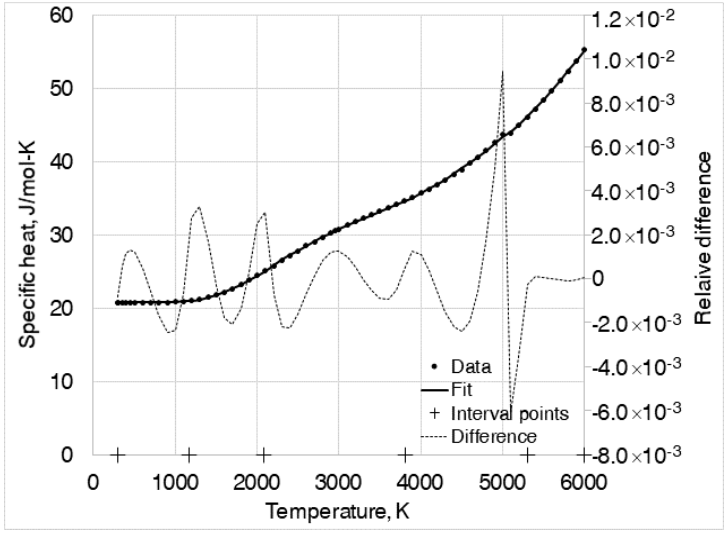


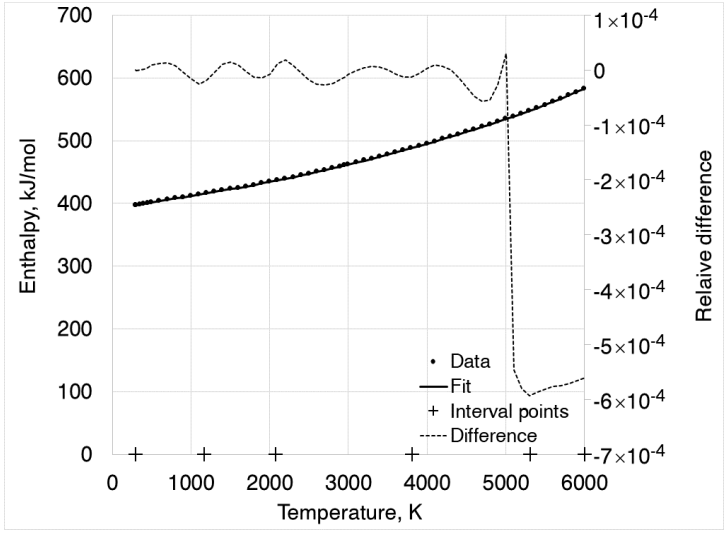


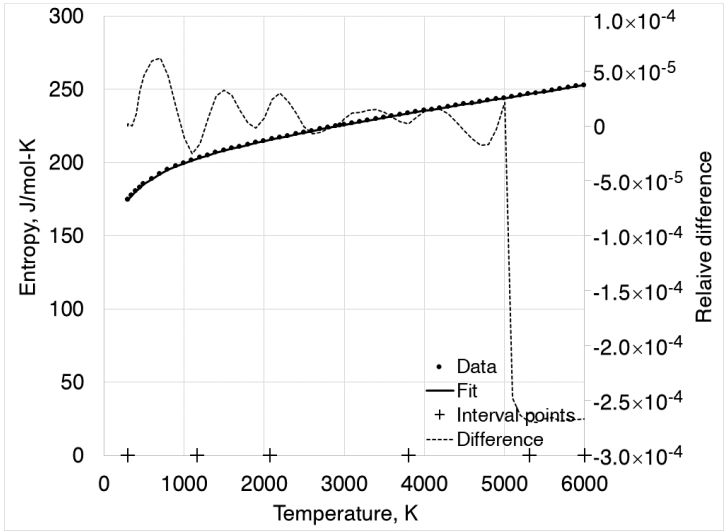


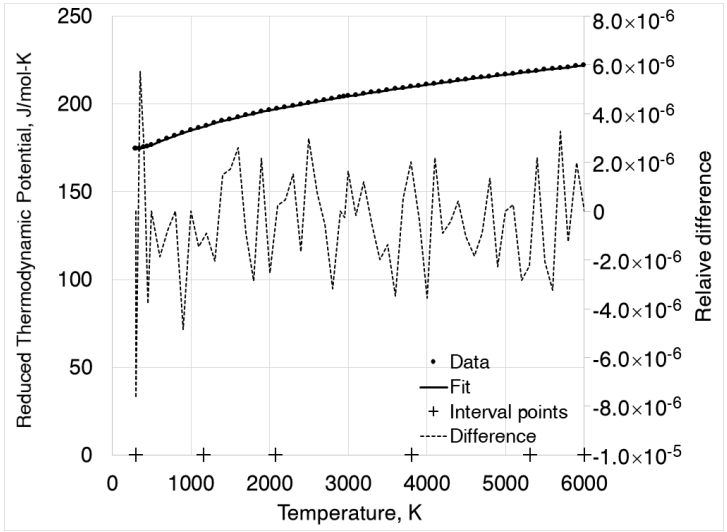
Chromium (Cr) – Cr1(g)

MW = 51.996 g/mol – $\Delta_f H^o$(298.15 K) = 397.48 kJ/mol – S^o(298.15 K) = 174.311 J/mol-K – p^o = 0.1 MPa



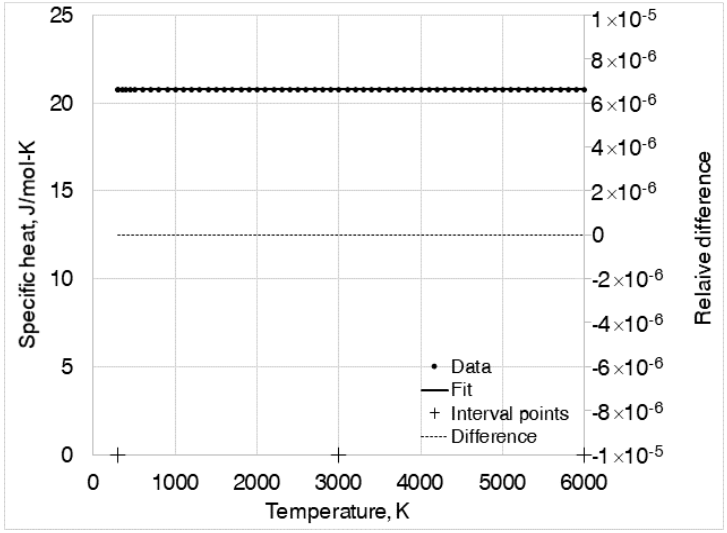


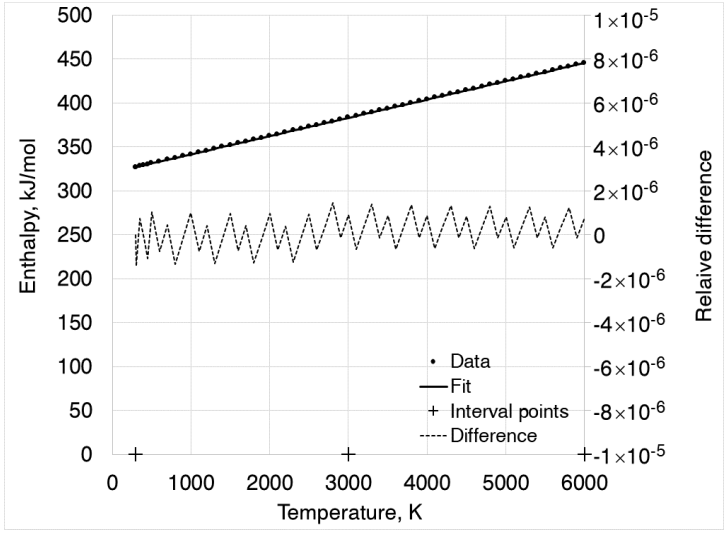


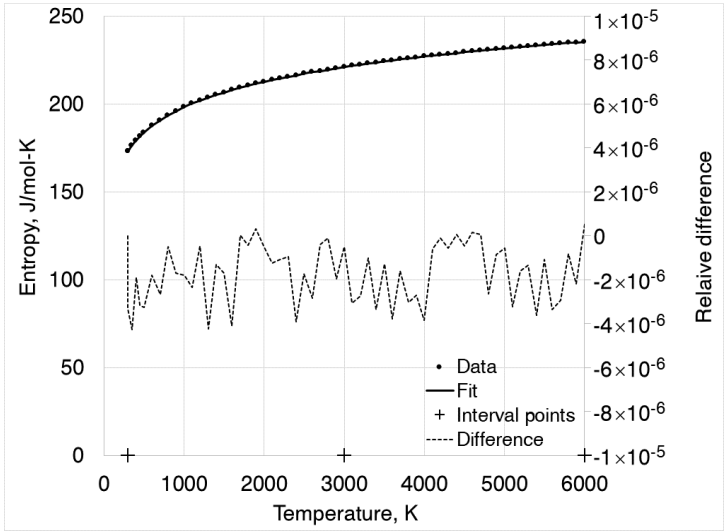


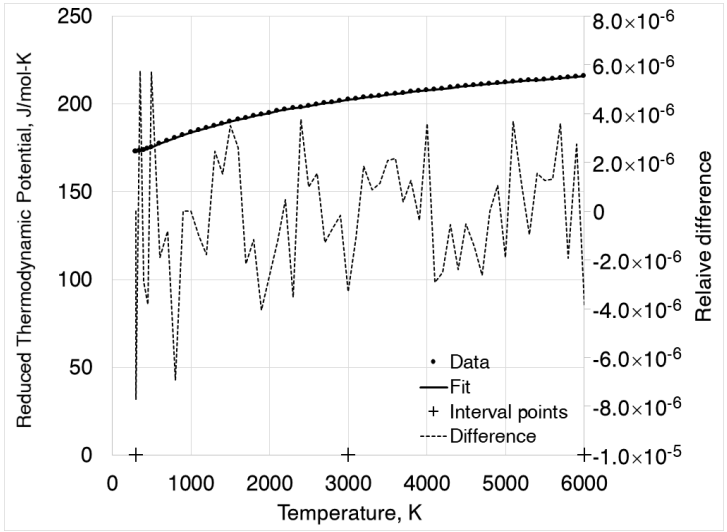
Chromium, Ion (Cr^-) – Cr1-(g)

MW = 51.997 g/mol – $\Delta_f H^o$(298.15 K) = 327.024 kJ/mol – S^o(298.15 K) = 173.03 J/mol-K – p^o = 0.1 MPa



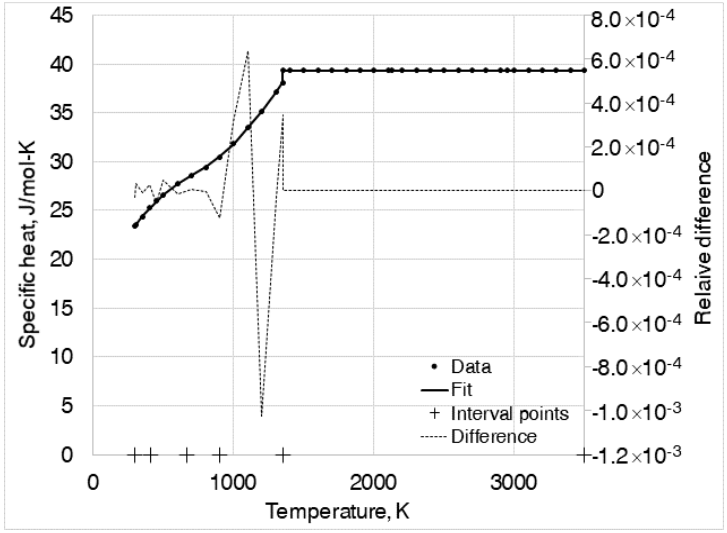


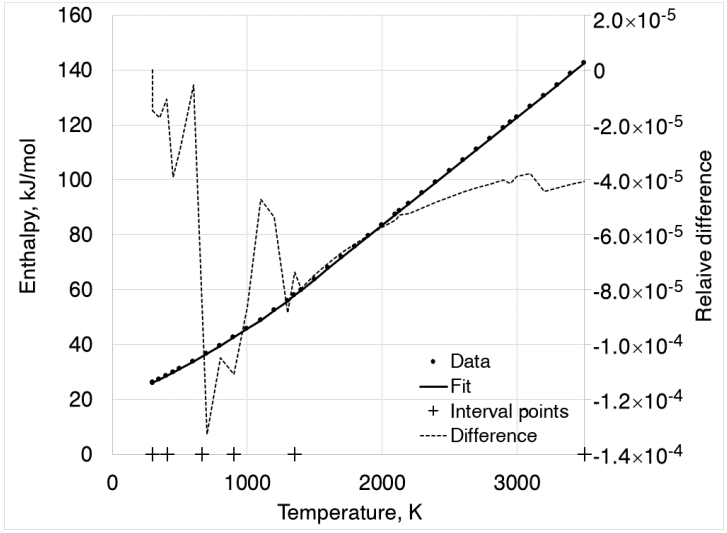


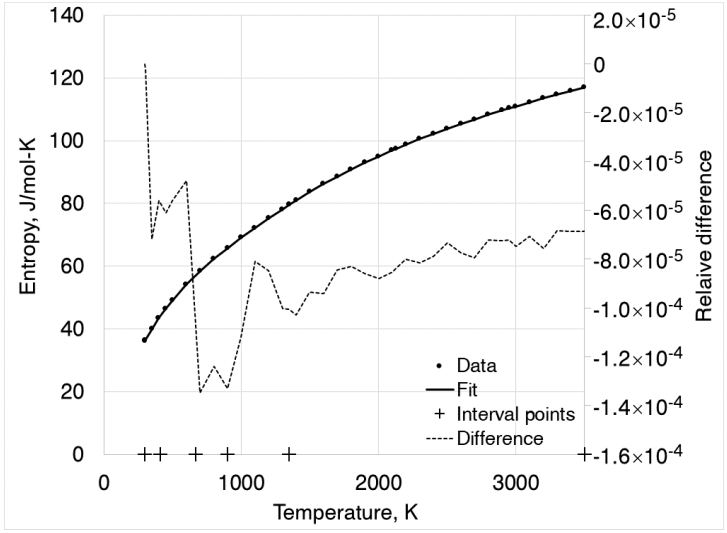


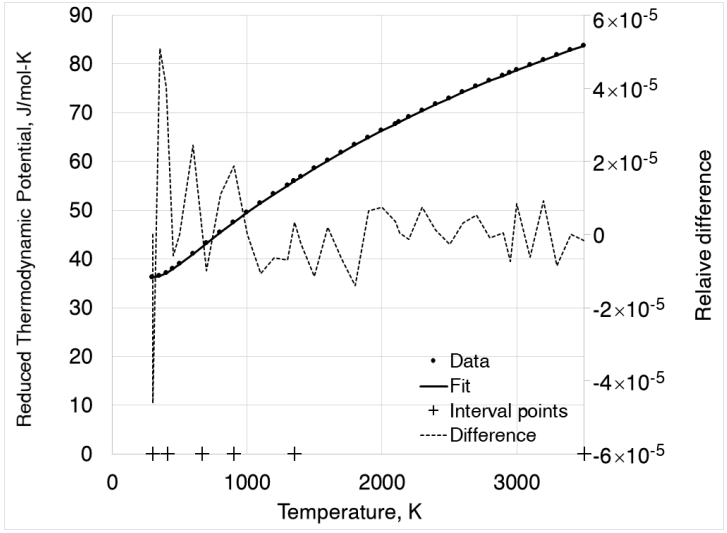
Chromium (Cr) – Cr1(l)

MW = 51.996 g/mol – $\Delta_f H^o$(298.15 K) = 26.068 kJ/mol – S^o(298.15 K) = 36.205 J/mol-K – p^o = 0.1 MPa



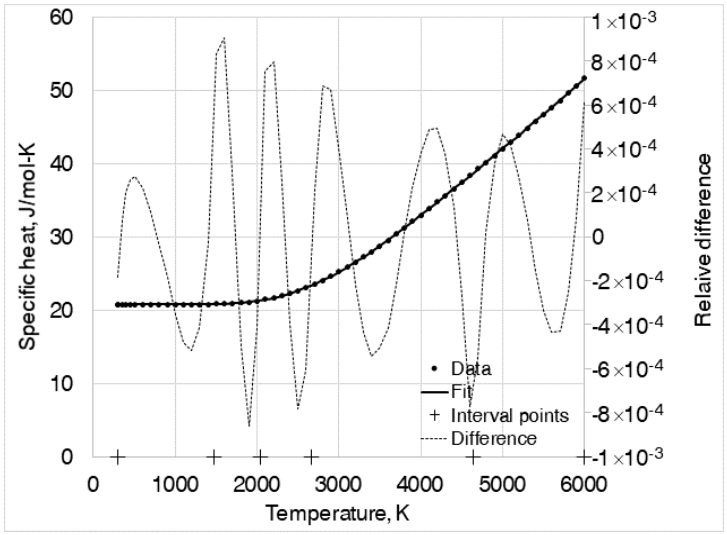


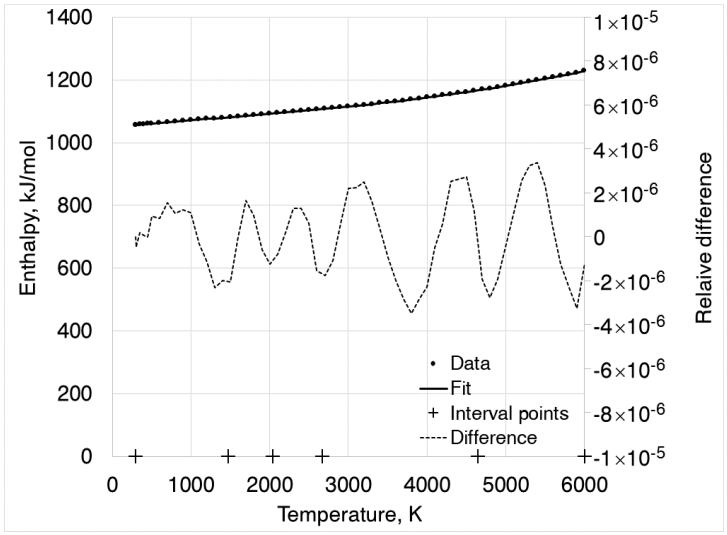


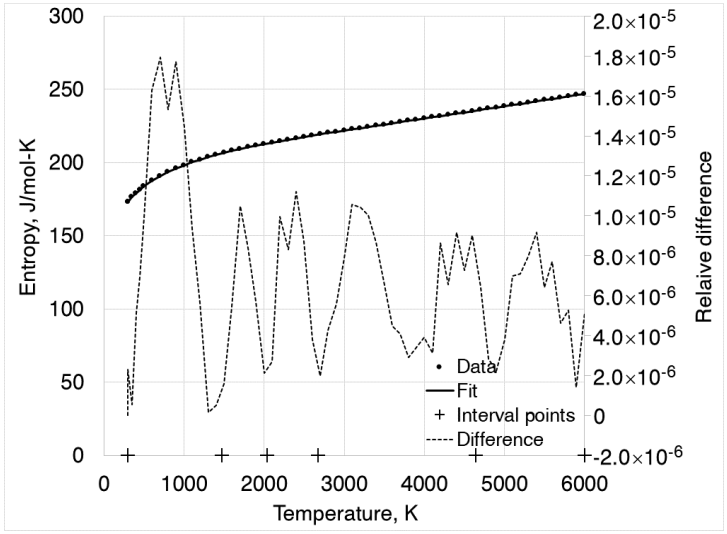


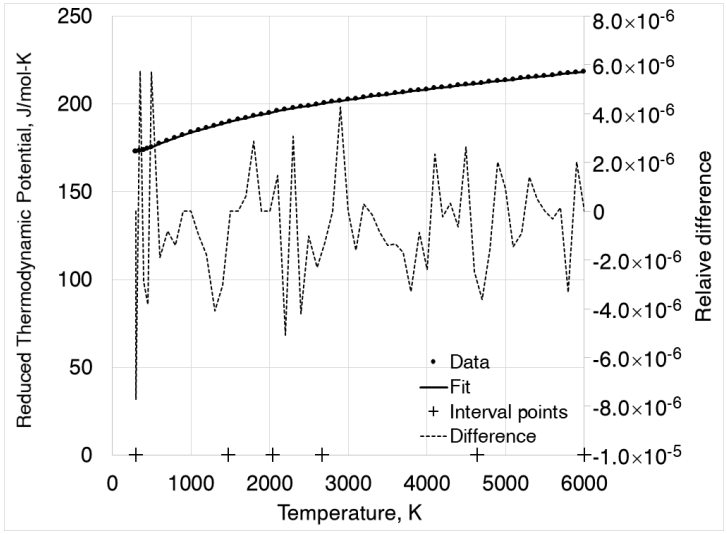
Chromium, Ion (Cr^{+}) – Cr1+(g)

MW = 51.995 g/mol – $\Delta_f H^o$(298.15 K) = 1056.48 kJ/mol – S^o(298.15 K) = 173.029 J/mol-K – p^o = 0.1 MPa



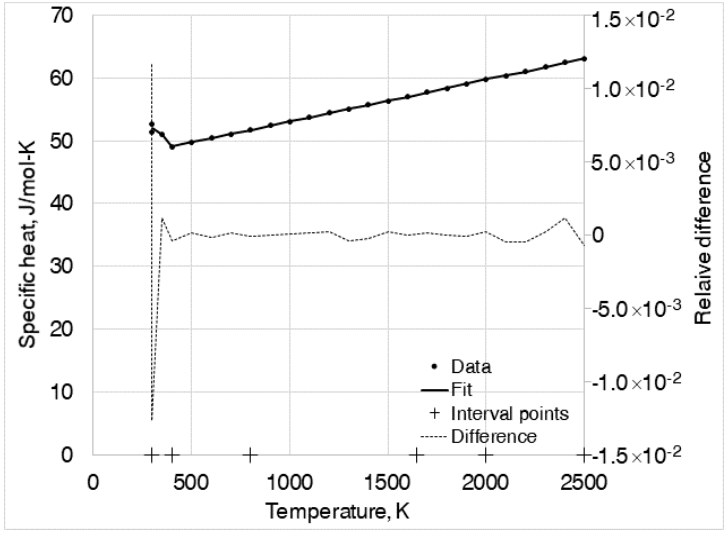


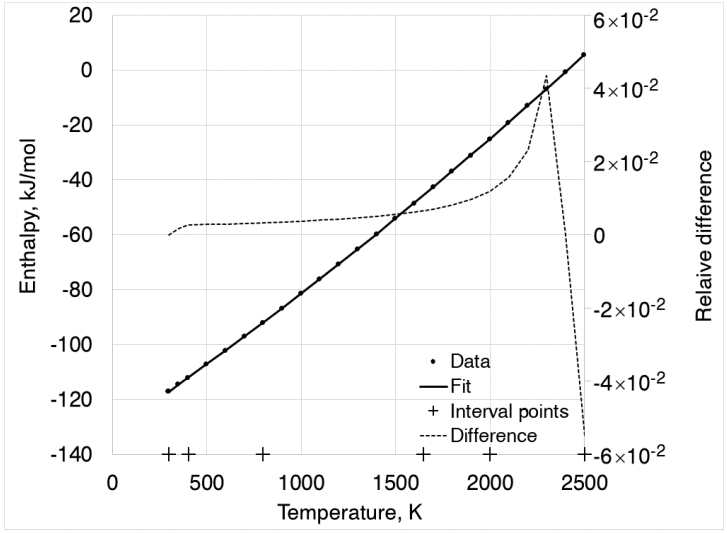


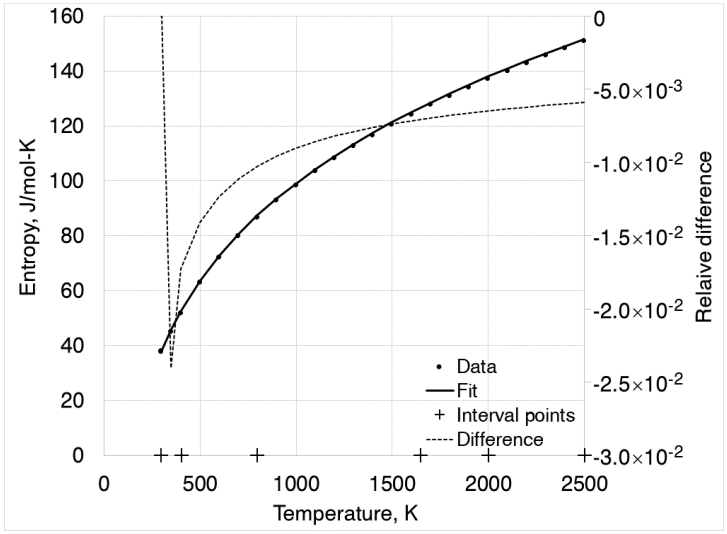


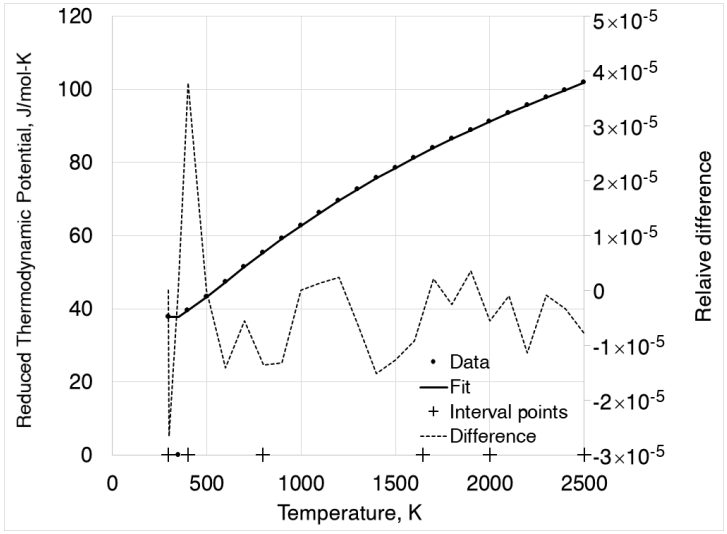
Chromium Nitride (CrN) – Cr1N1(cr)

MW = 66.003 g/mol – $\Delta_f H^o$(298.15 K) = -117.152 kJ/mol – S^o(298.15 K) = 37.711 J/mol-K – p^o = 0.1 MPa



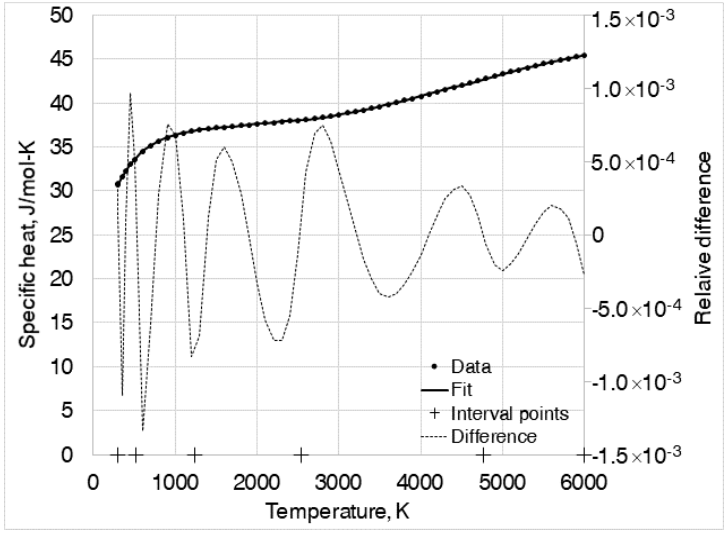


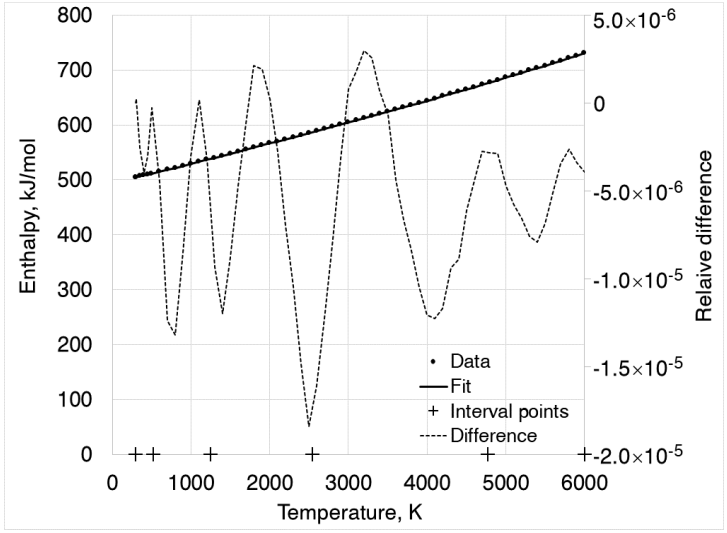


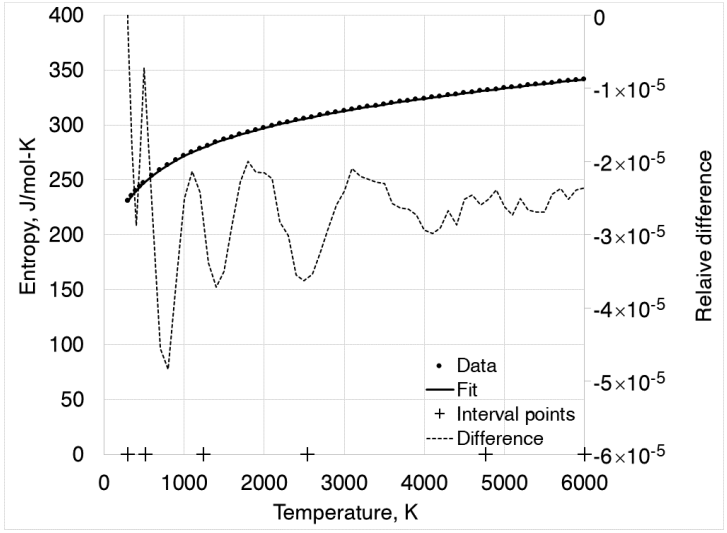


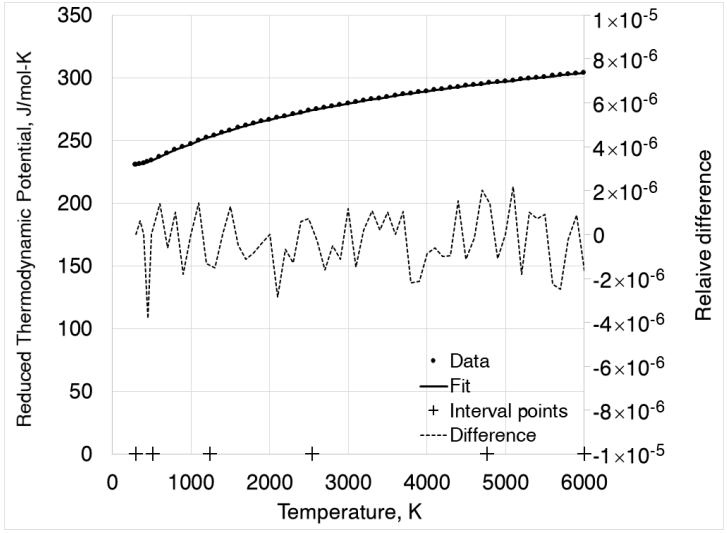
Chromium Nitride (CrN) – Cr1N1(g)

MW = 66.003 g/mol – $\Delta_f H^o$(298.15 K) = 505.009 kJ/mol – S^o(298.15 K) = 230.553 J/mol-K – p^o = 0.1 MPa



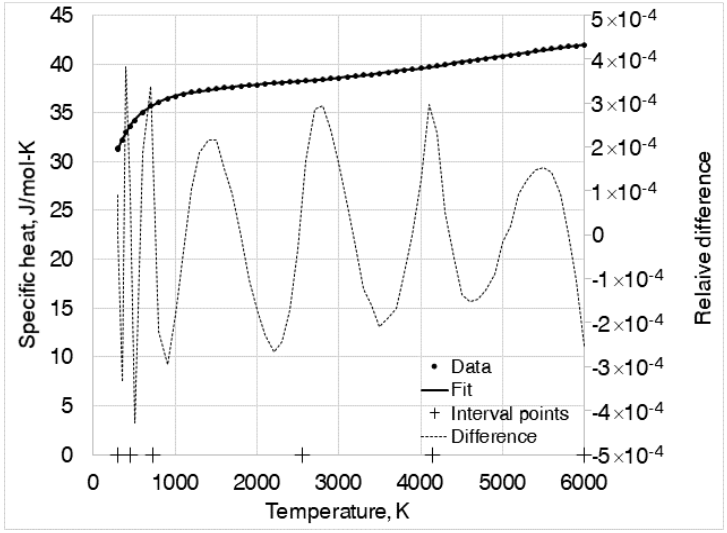


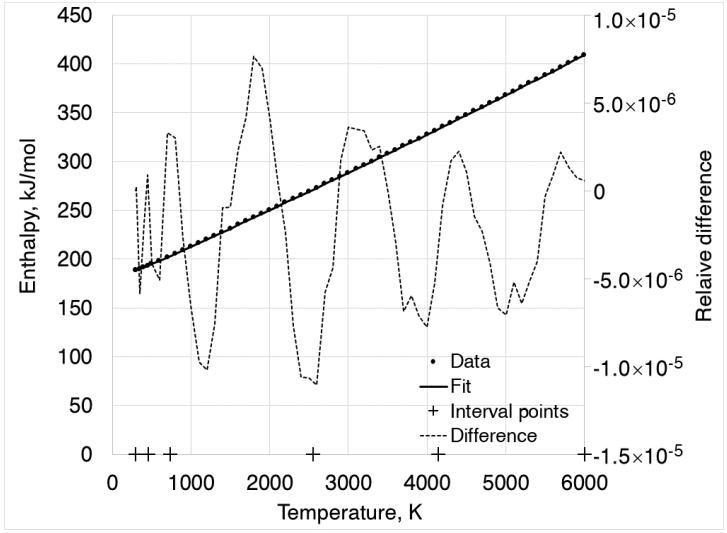


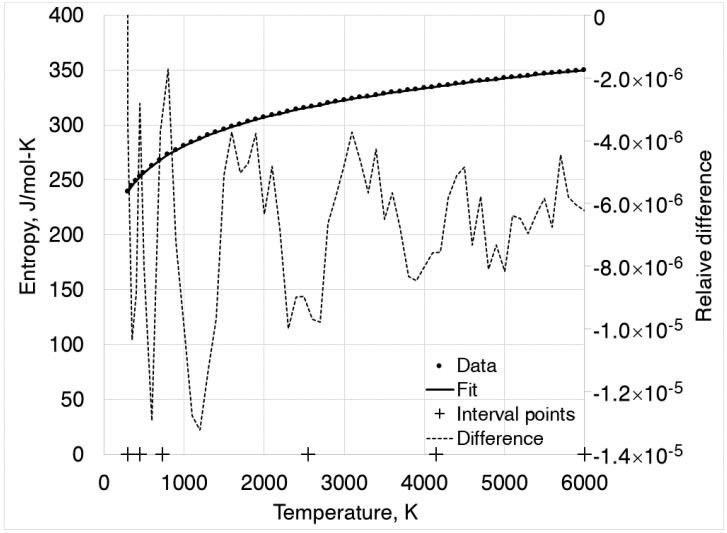


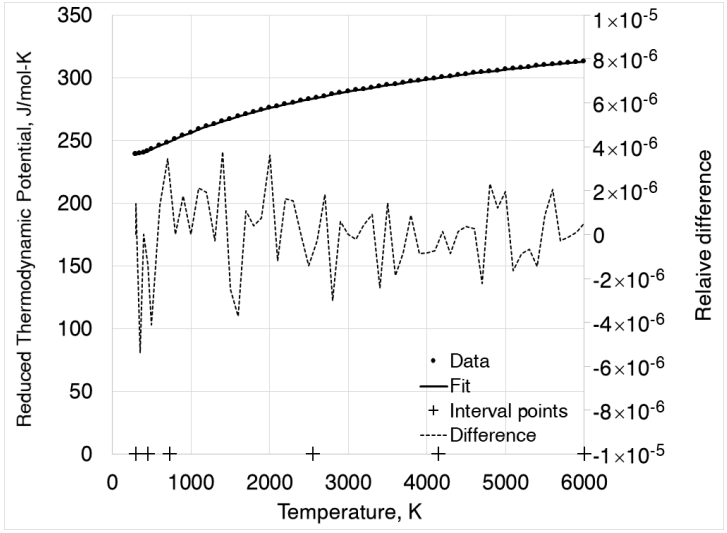
Chromium Oxide (CrO) – Cr1O1(g)

MW = 67.995 g/mol – $\Delta_f H^o$(298.15 K) = 188.28 kJ/mol – S^o(298.15 K) = 239.268 J/mol-K – p^o = 0.1 MPa



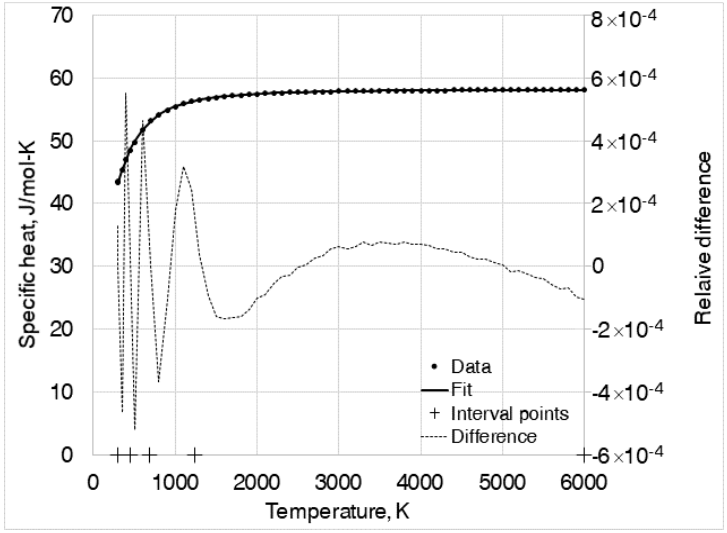


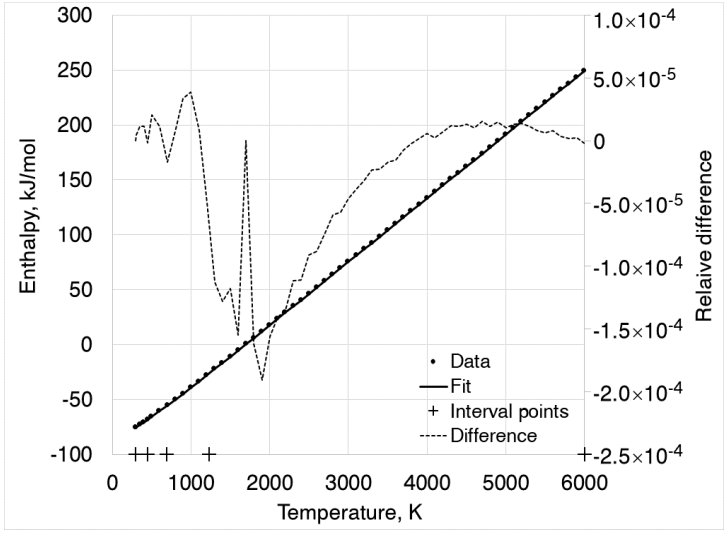


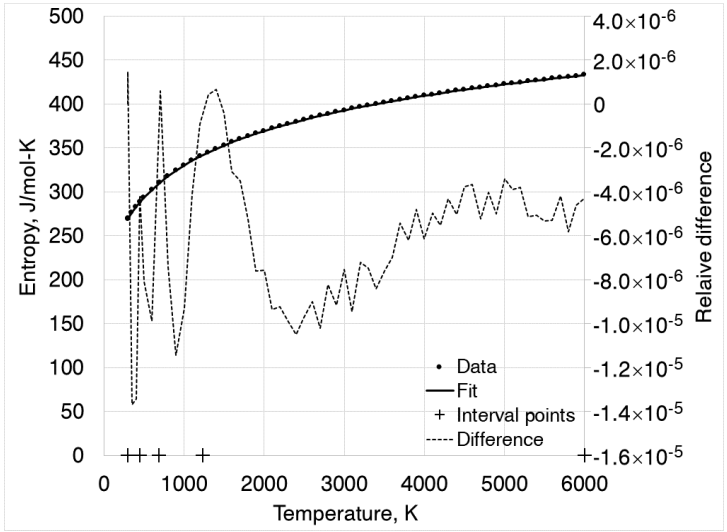


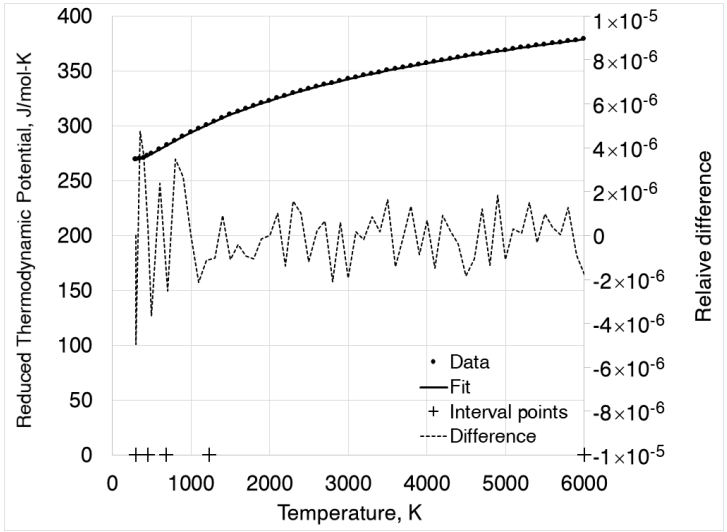
Chromium Oxide (CrO_2) – Cr1O2(g)

MW = 83.995 g/mol – $\Delta_f H^o$(298.15 K) = -75.312 kJ/mol – S^o(298.15 K) = 269.241 J/mol-K – p^o = 0.1 MPa



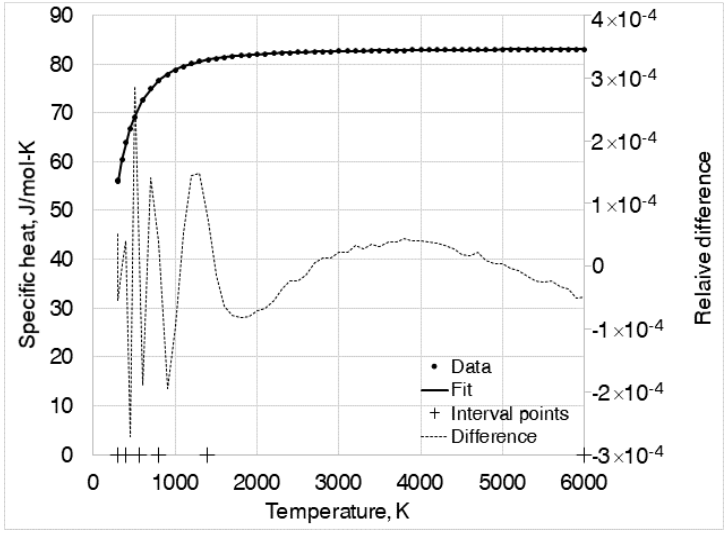


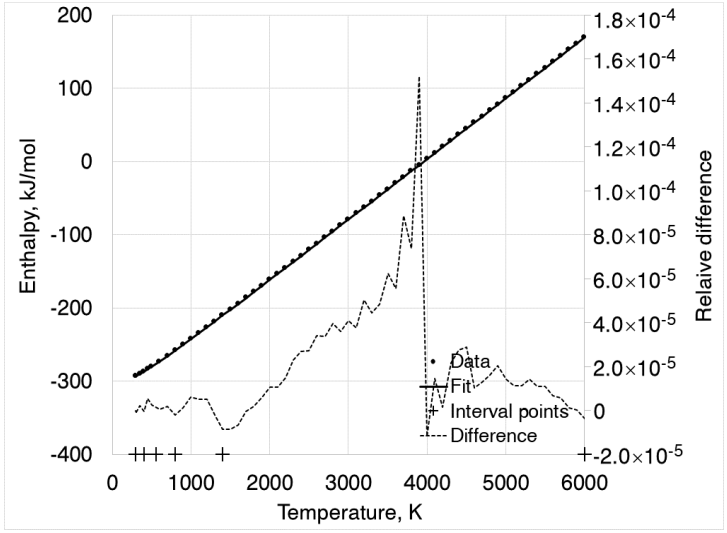


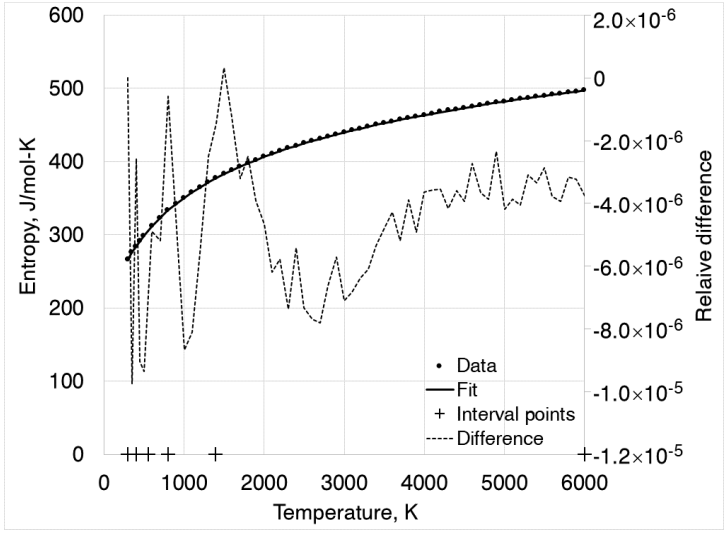


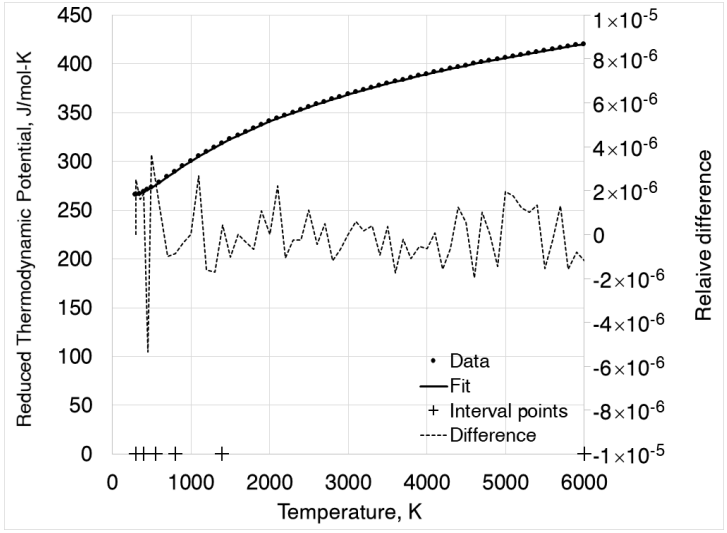
Chromium Oxide (CrO_3) – Cr1O3(g)

MW = 99.994 g/mol – $\Delta_f H^o$(298.15 K) = -292.88 kJ/mol – S^o(298.15 K) = 266.178 J/mol-K – p^o = 0.1 MPa



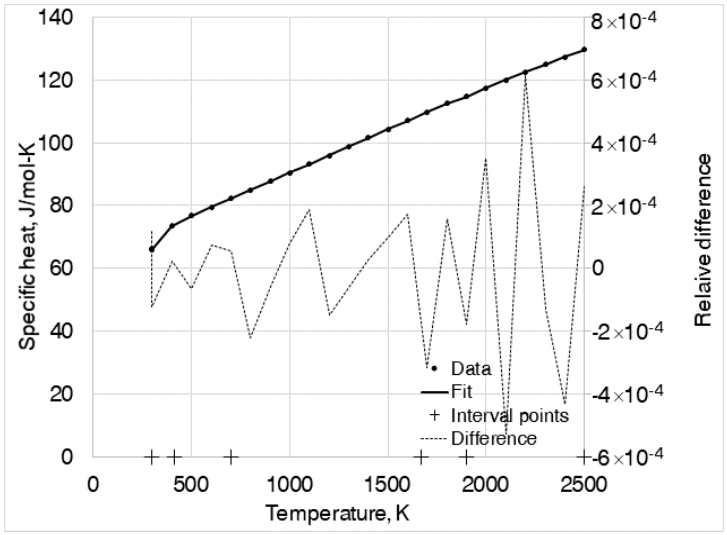


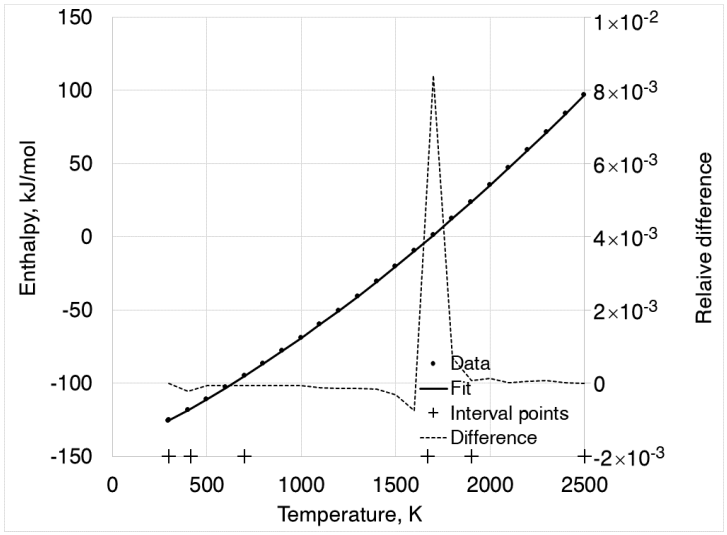


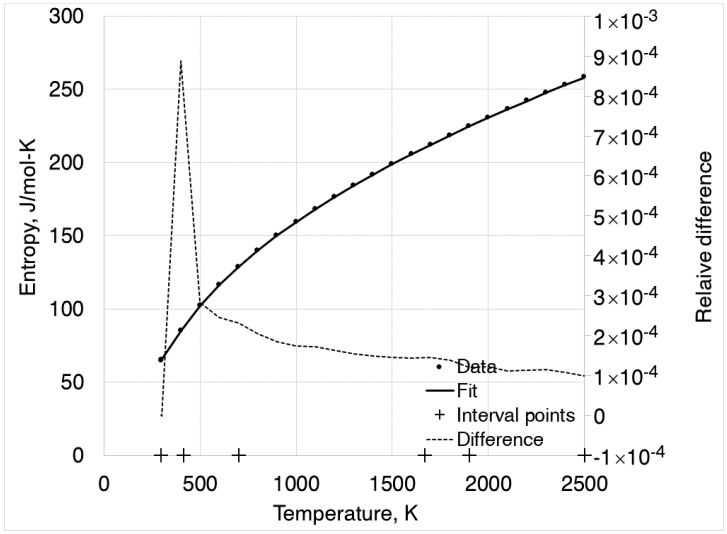


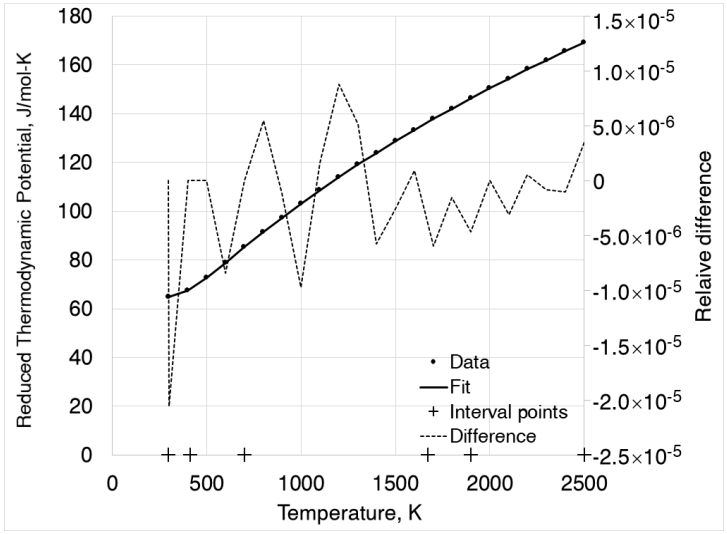
Chromium Nitride (Cr_2N) – Cr2N1(cr)

MW = 117.999 g/mol – $\Delta_f H^o$(298.15 K) = -125.52 kJ/mol – S^o(298.15 K) = 64.852 J/mol-K – p^o = 0.1 MPa



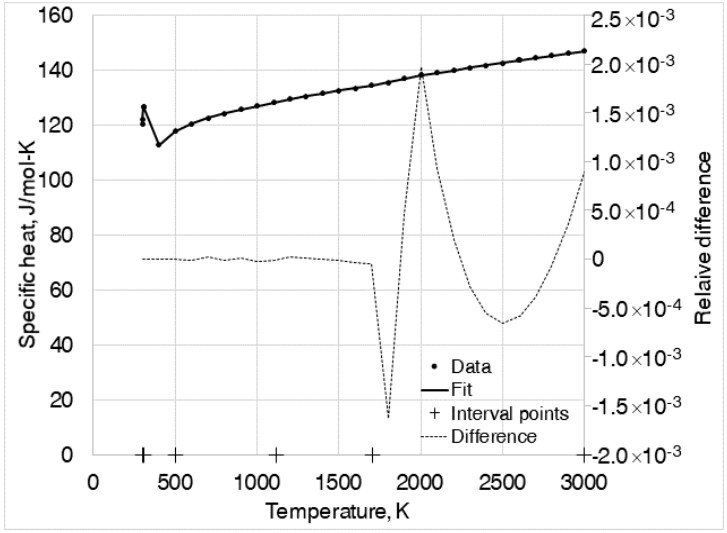


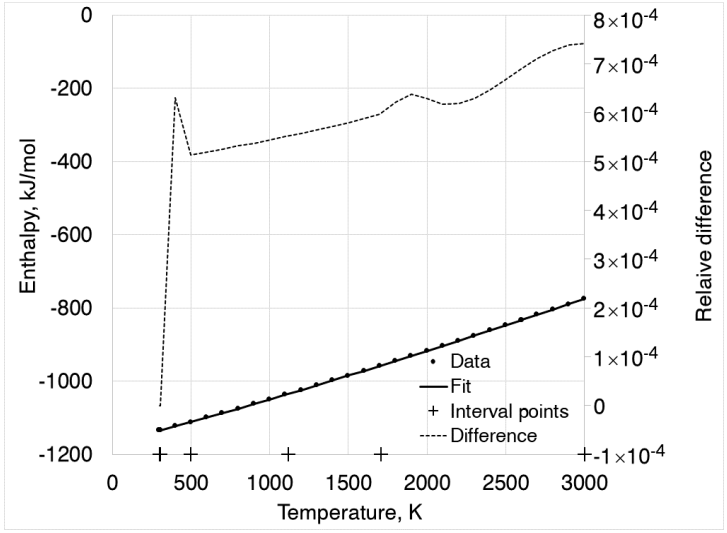


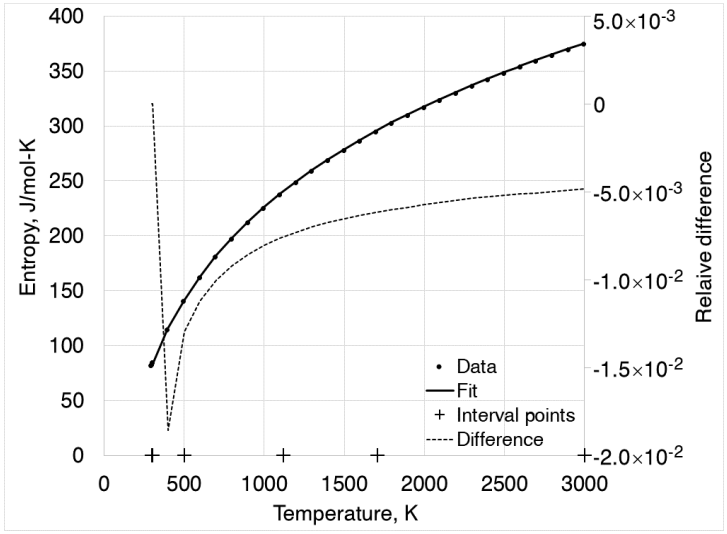


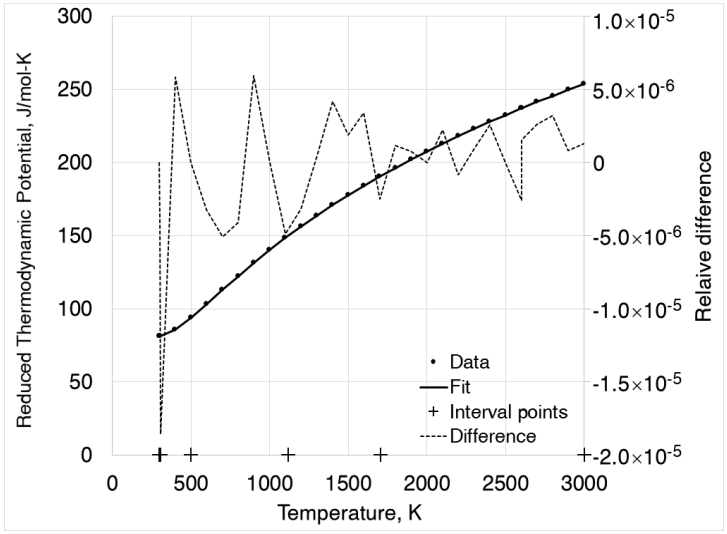
Chromium Oxide (Cr_2O_3) – Cr2O3(cr)

MW = 151.99 g/mol – $\Delta_f H^o$(298.15 K) = -1134.701 kJ/mol – S^o(298.15 K) = 81.154 J/mol-K – p^o = 0.1 MPa



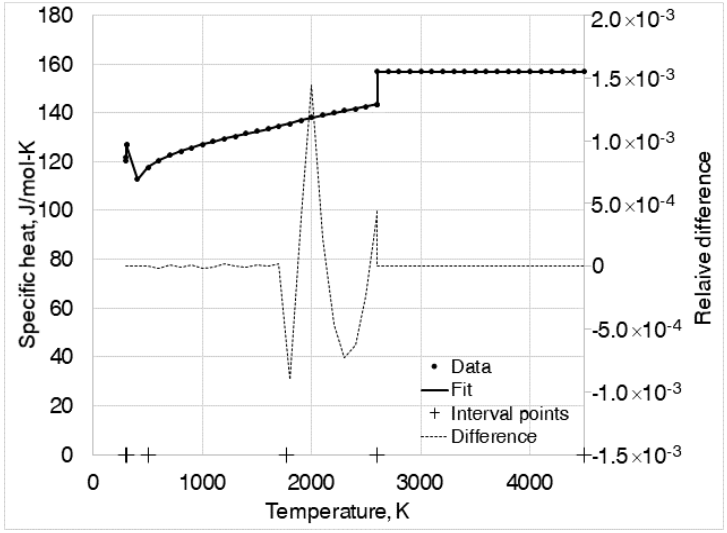


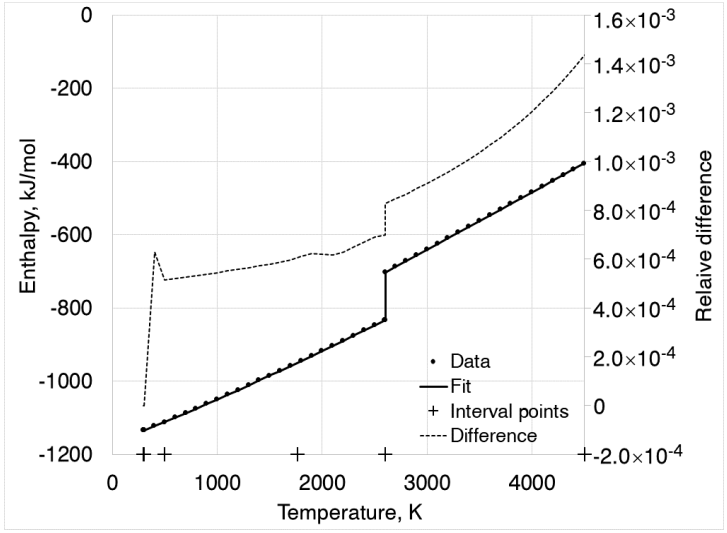


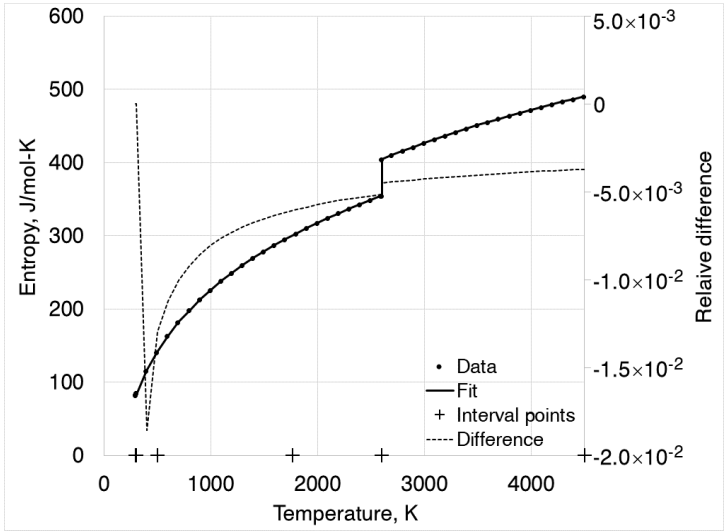


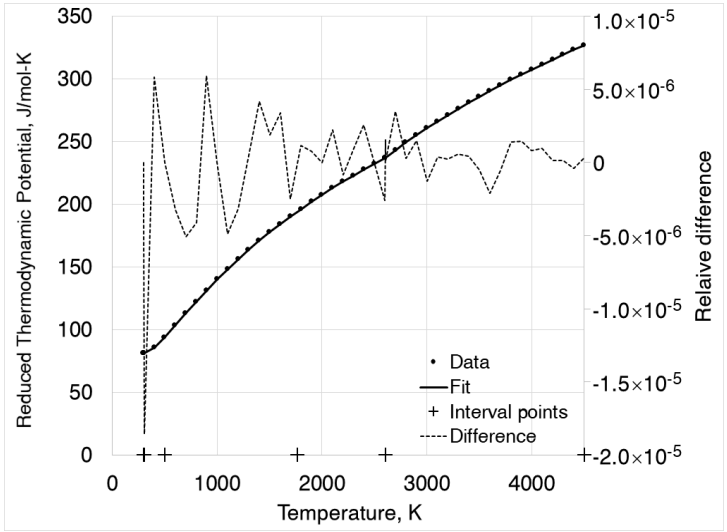
Chromium Oxide (Cr_2O_3) – Cr2O3(cr,l)

MW = 151.99 g/mol – $\Delta_f H^o$(298.15 K) = -1134.701 kJ/mol – S^o(298.15 K) = 81.154 J/mol-K – p^o = 0.1 MPa



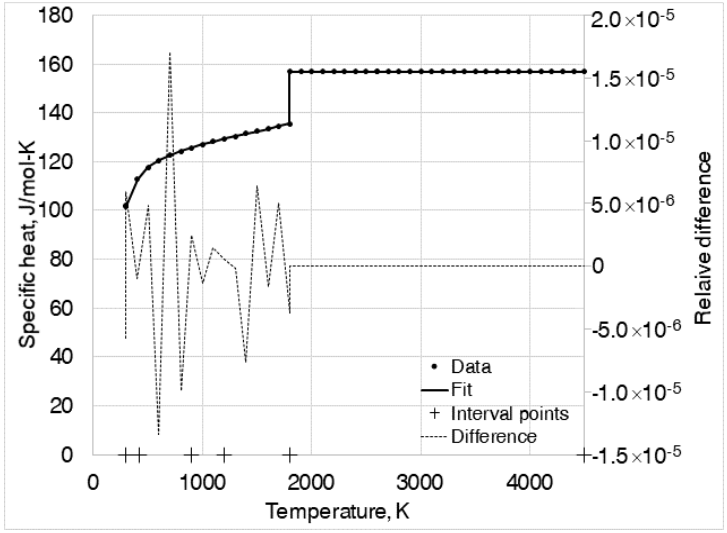


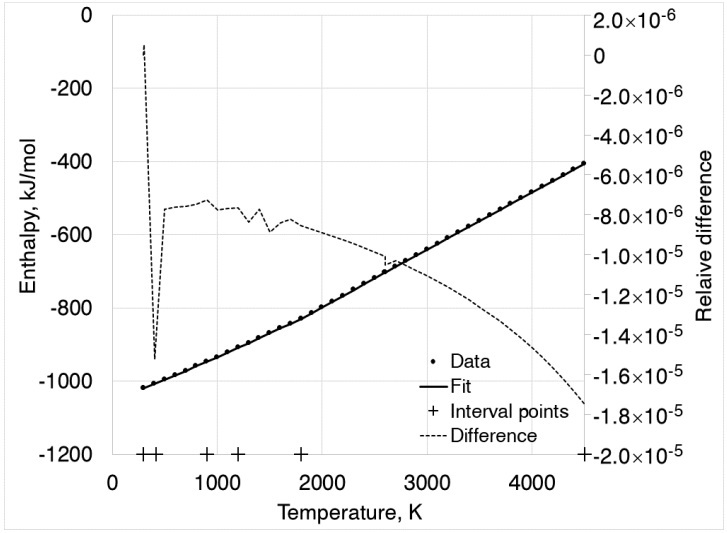


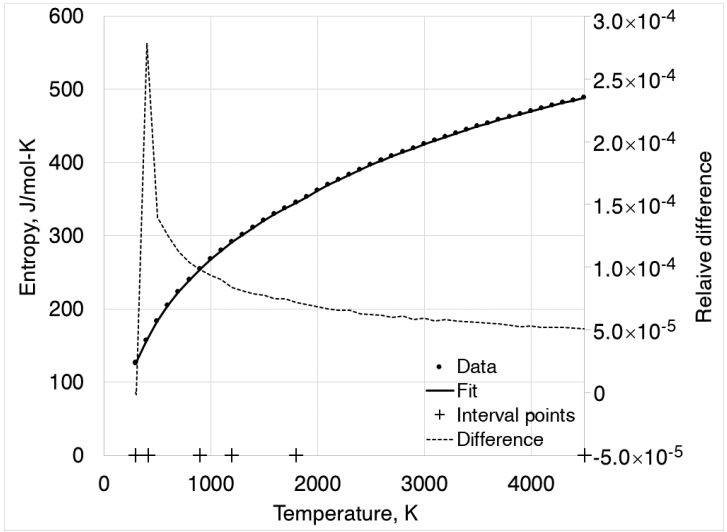


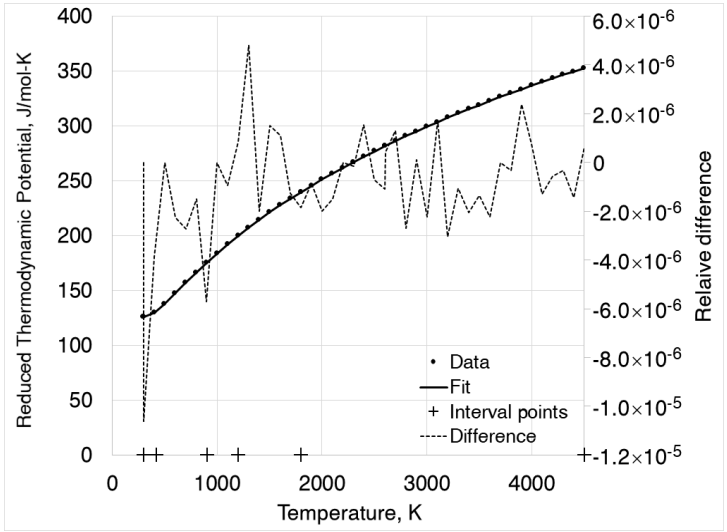
Chromium Oxide (Cr_2O_3) – Cr2O3(l)

MW = 151.99 g/mol – $\Delta_f H^o$(298.15 K) = -1018.381 kJ/mol – S^o(298.15 K) = 125.595 J/mol-K – p^o = 0.1 MPa



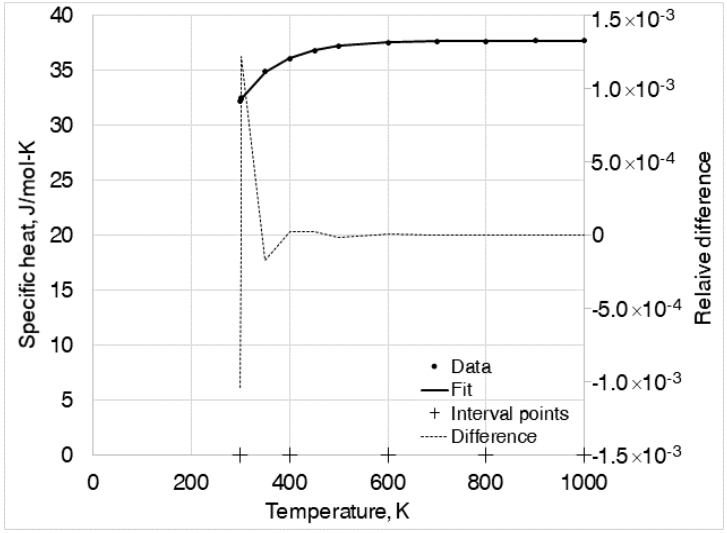


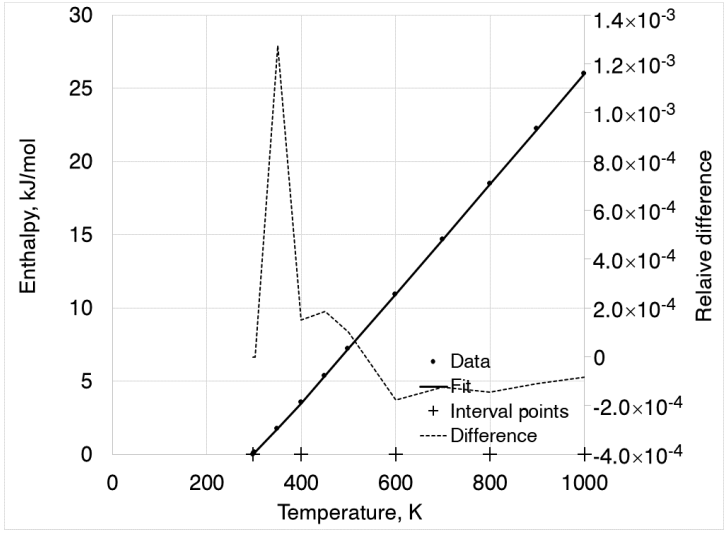


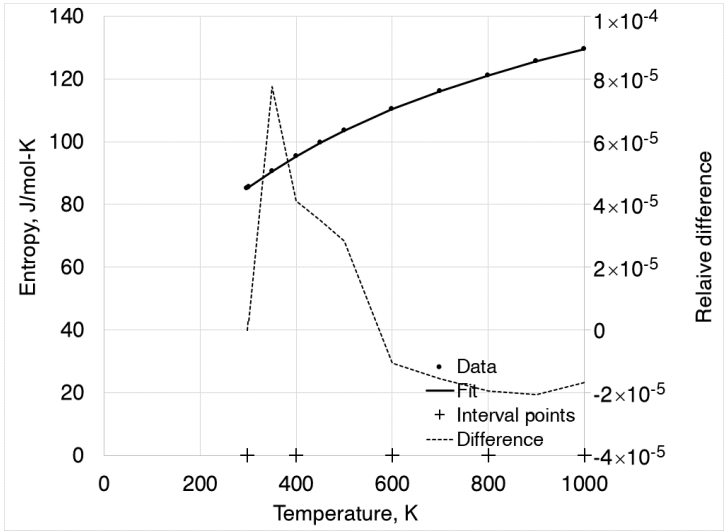


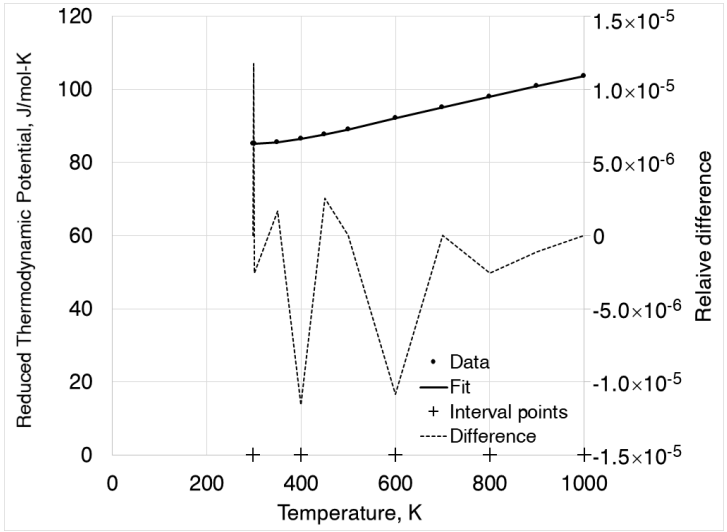
Cesium (Cs) – Cs1(cr)

MW = 132.9054 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 85.147 J/mol-K – p^o = 0.1 MPa



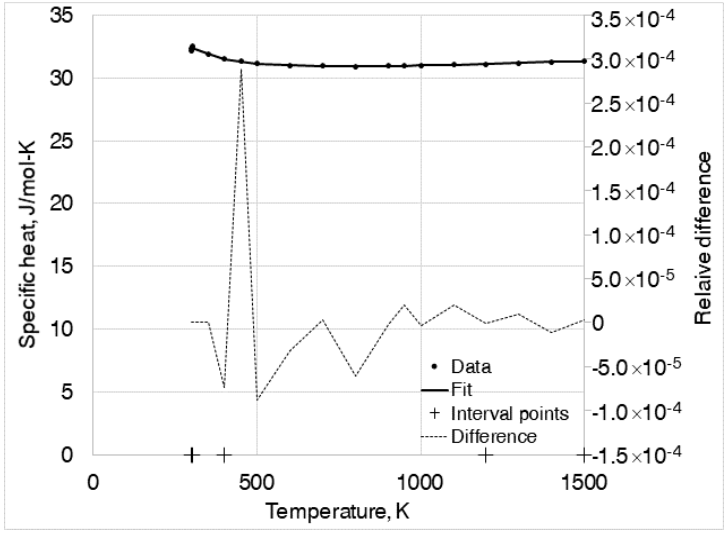


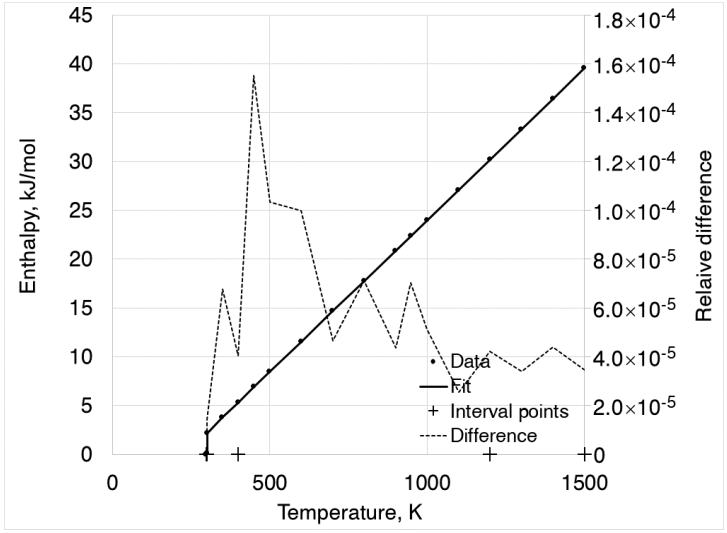


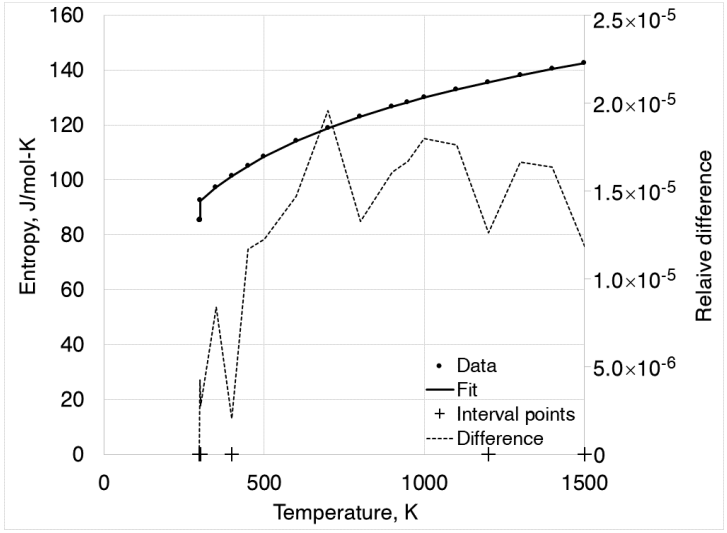


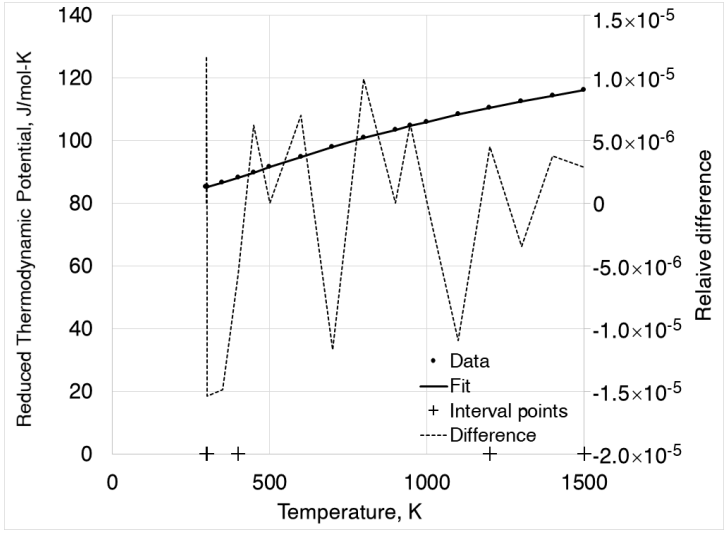
Cesium (Cs) – Cs1(cr,l)

MW = 132.9054 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 85.147 J/mol-K – p^o = 0.1 MPa



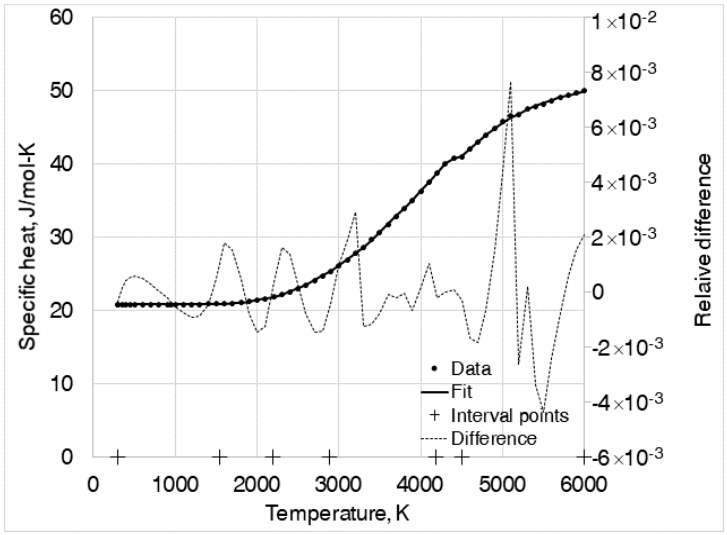


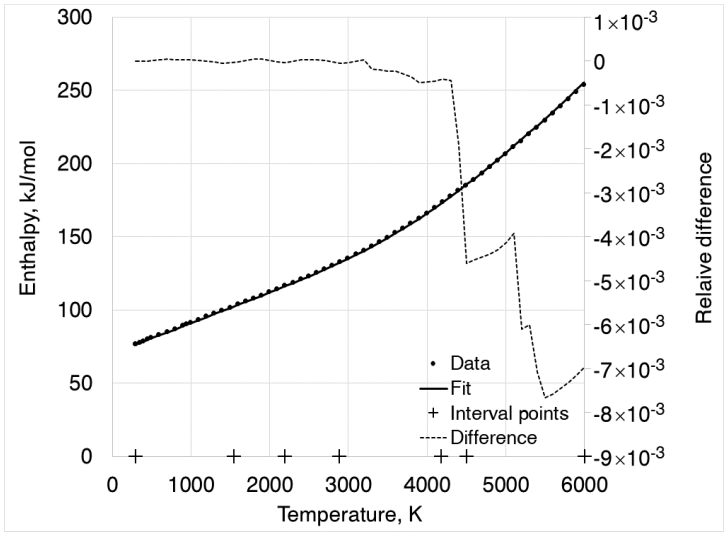


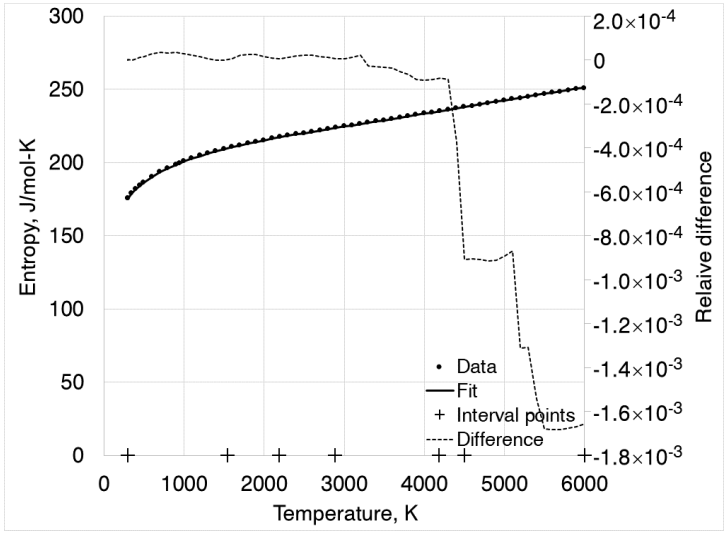


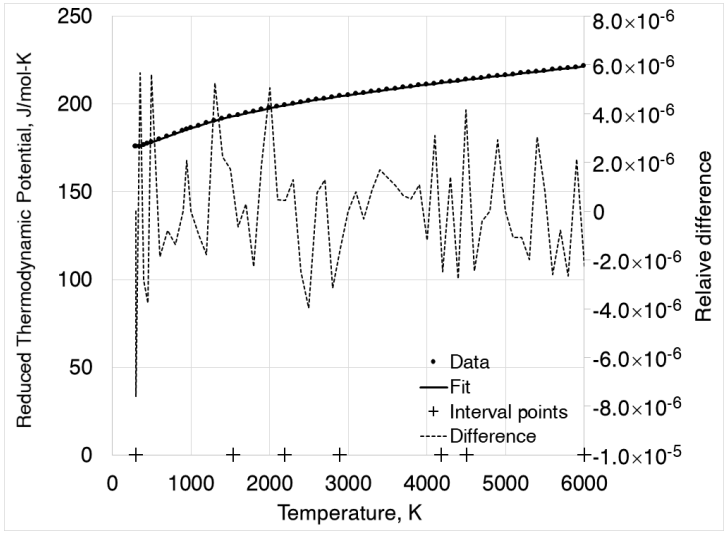
Cesium (Cs) – Cs1(g)

MW = 132.9054 g/mol – $\Delta_f H^o$(298.15 K) = 76.5 kJ/mol – S^o(298.15 K) = 175.599 J/mol-K – p^o = 0.1 MPa



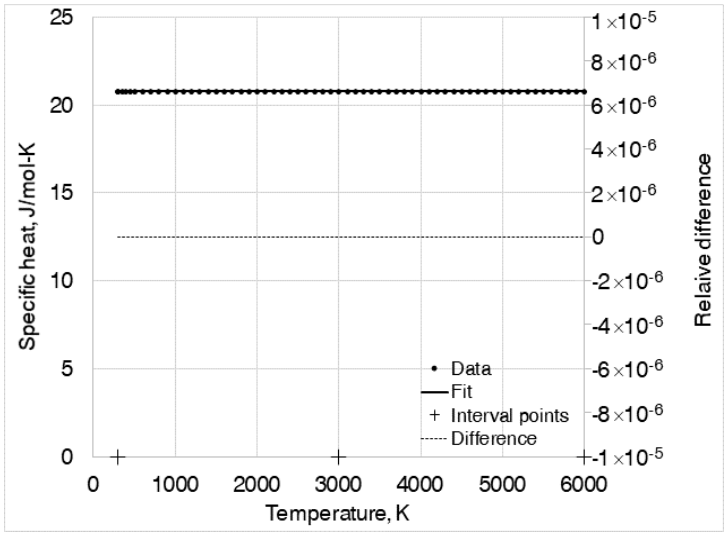


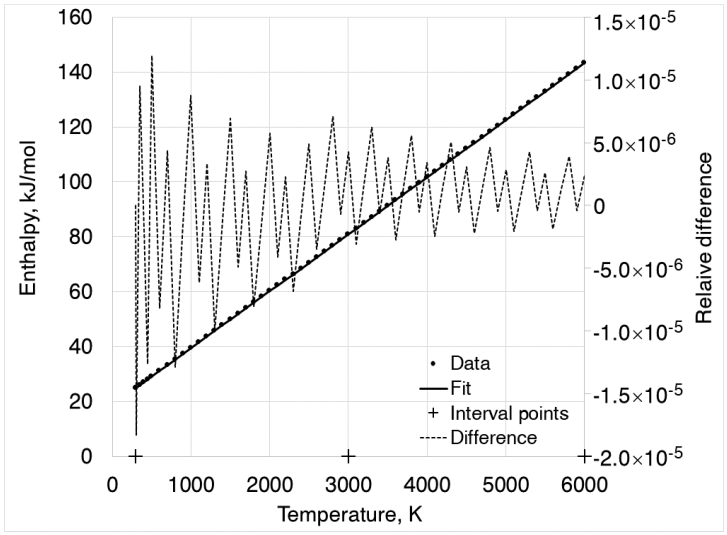


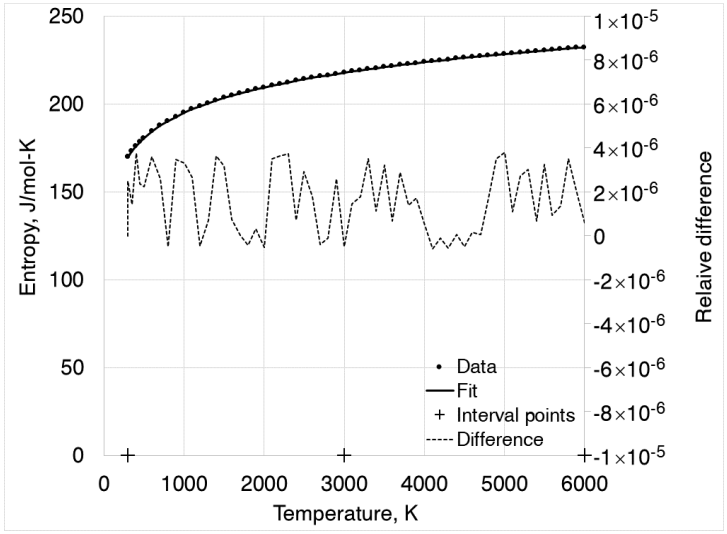


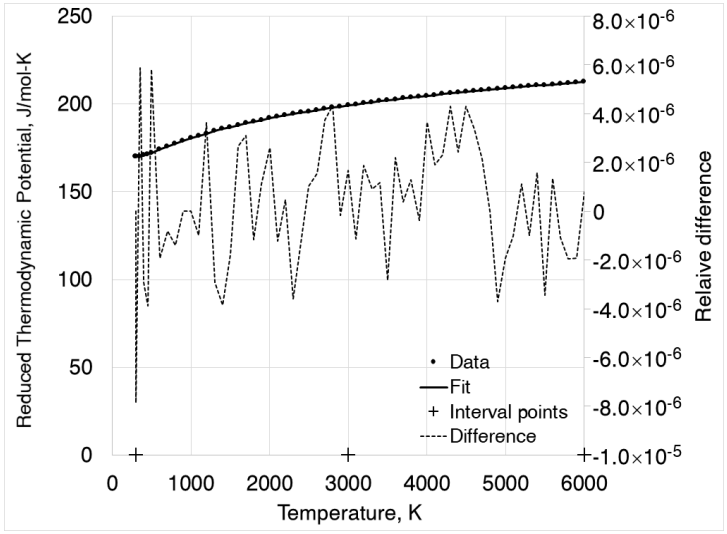
Cesium, Ion (Cs^-) – Cs1-(g)

MW = 132.9059 g/mol – $\Delta_f H^o$(298.15 K) = 24.798 kJ/mol – S^o(298.15 K) = 169.836 J/mol-K – p^o = 0.1 MPa



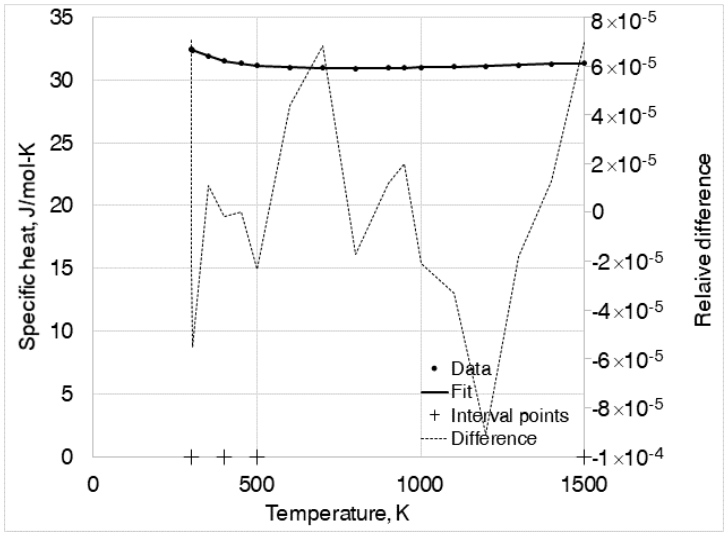


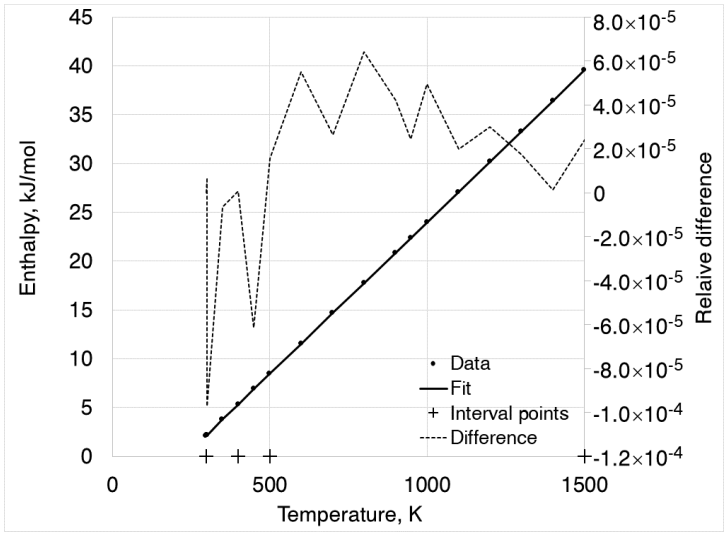


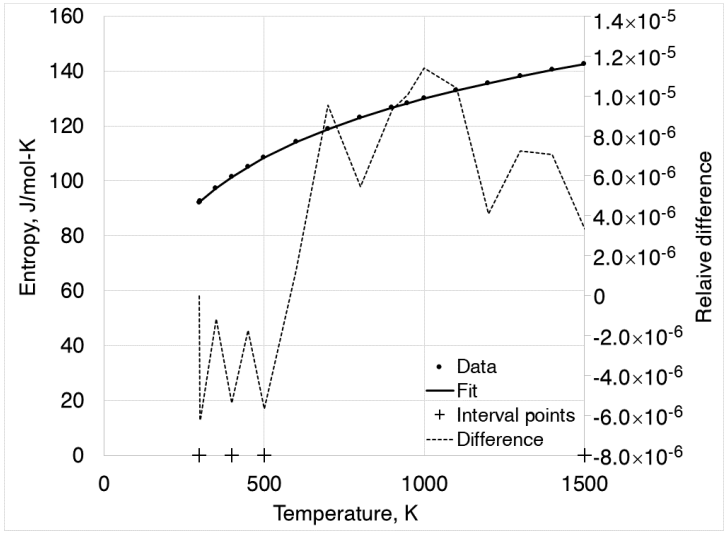


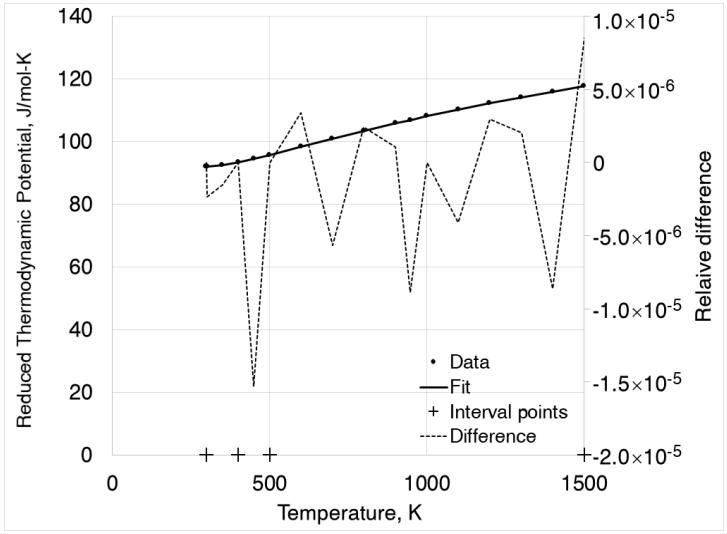
Cesium (Cs) – Cs1(l)

MW = 132.9054 g/mol – $\Delta_f H^o$(298.15 K) = 2.087 kJ/mol – S^o(298.15 K) = 92.069 J/mol-K – p^o = 0.1 MPa



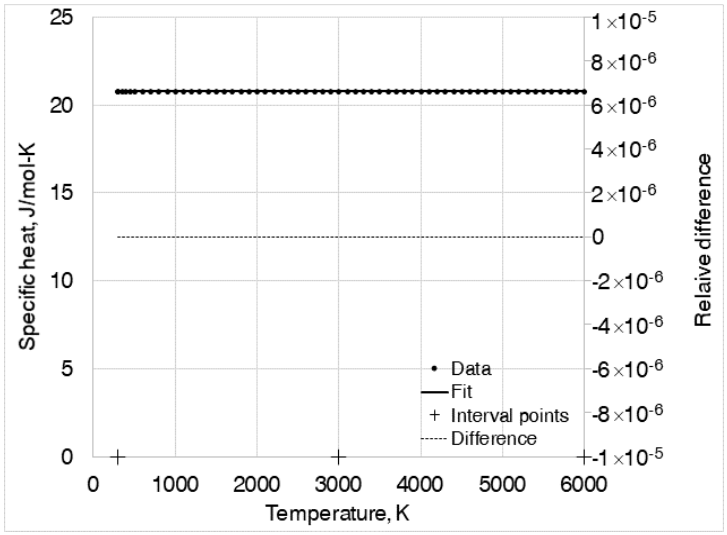


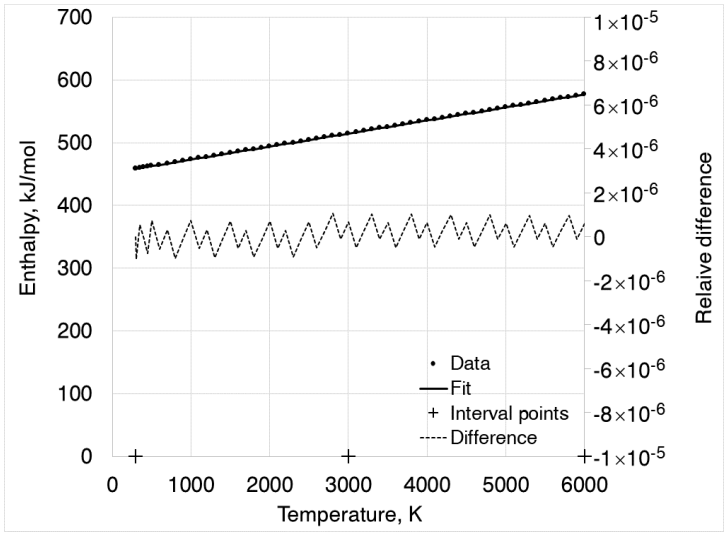


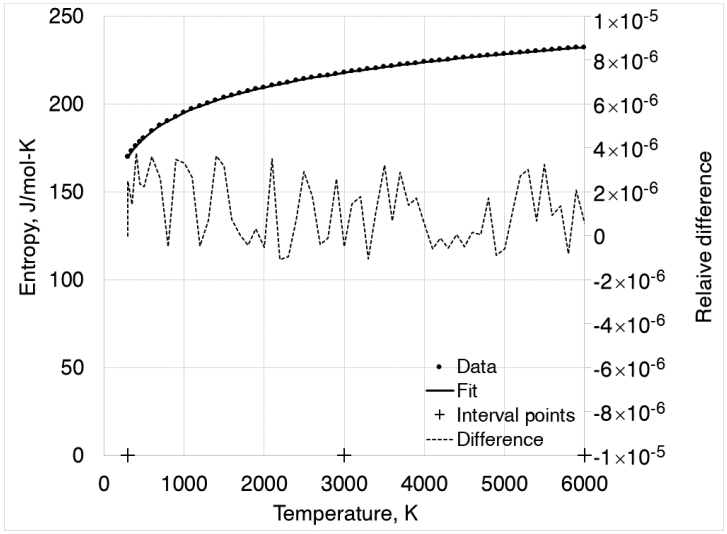


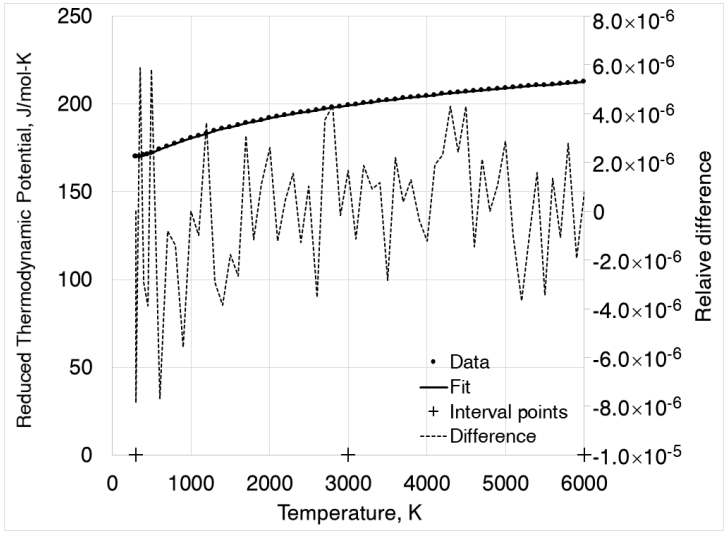
Cesium, Ion (Cs^+) – Cs1+(g)

MW = 132.9049 g/mol – $\Delta_f H^o$(298.15 K) = 458.402 kJ/mol – S^o(298.15 K) = 169.836 J/mol-K – p^o = 0.1 MPa



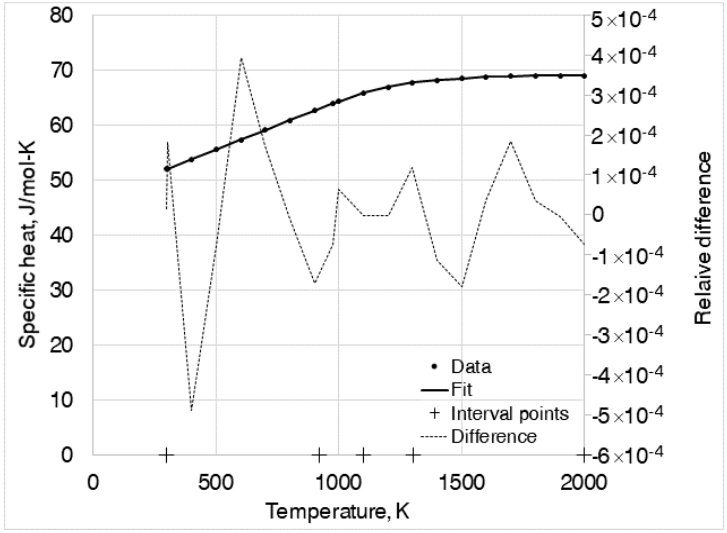


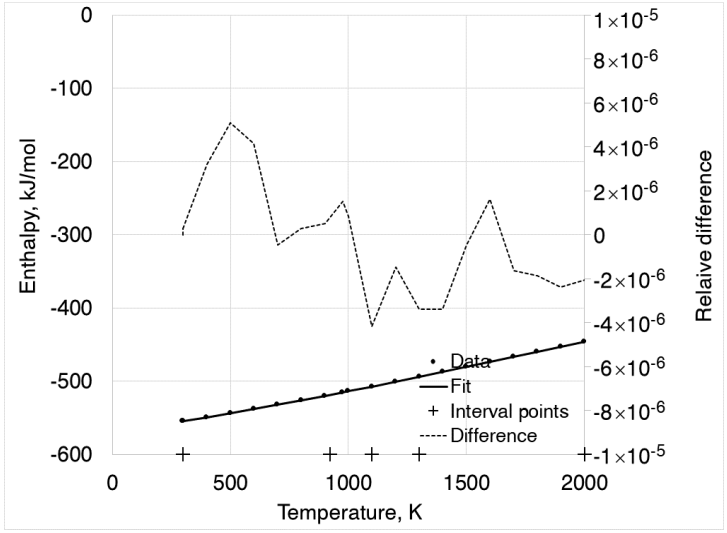


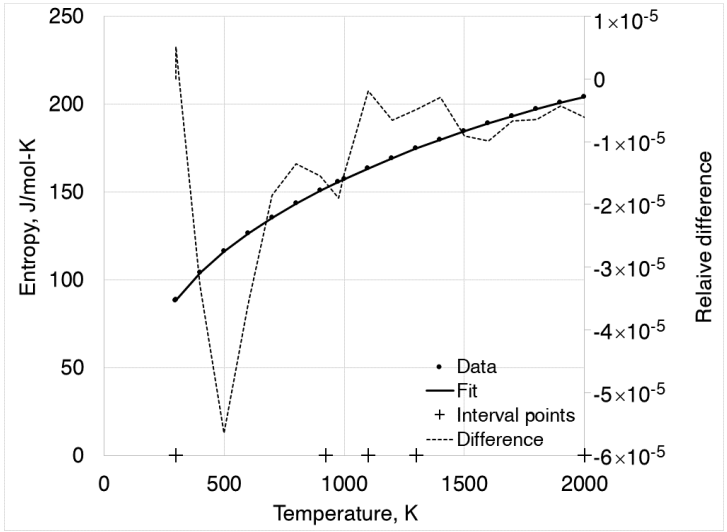


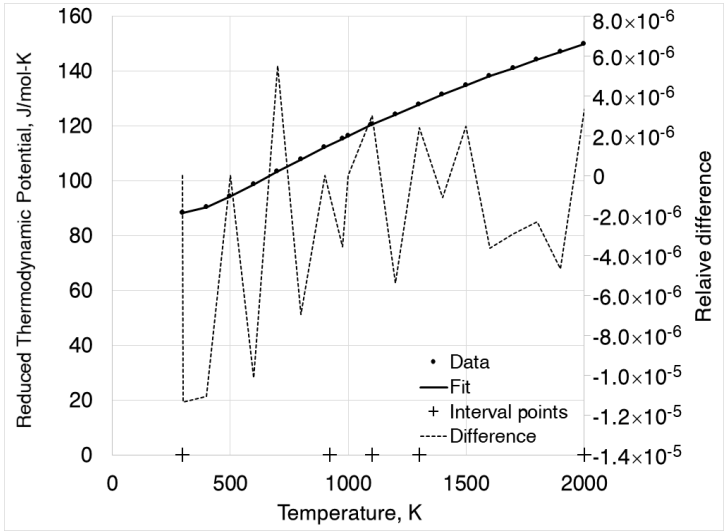
Cesium Fluoride (CsF) – Cs1F1(cr)

MW = 151.9038 g/mol – $\Delta_f H^o$(298.15 K) = -554.673 kJ/mol – S^o(298.15 K) = 88.282 J/mol-K – p^o = 0.1 MPa



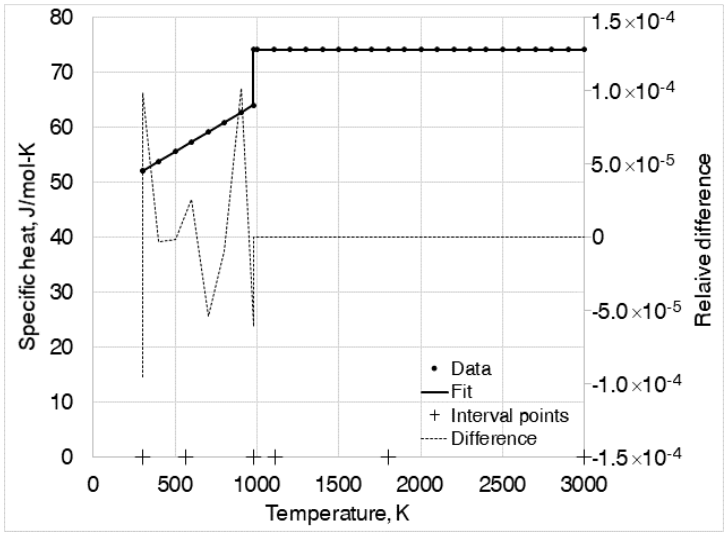


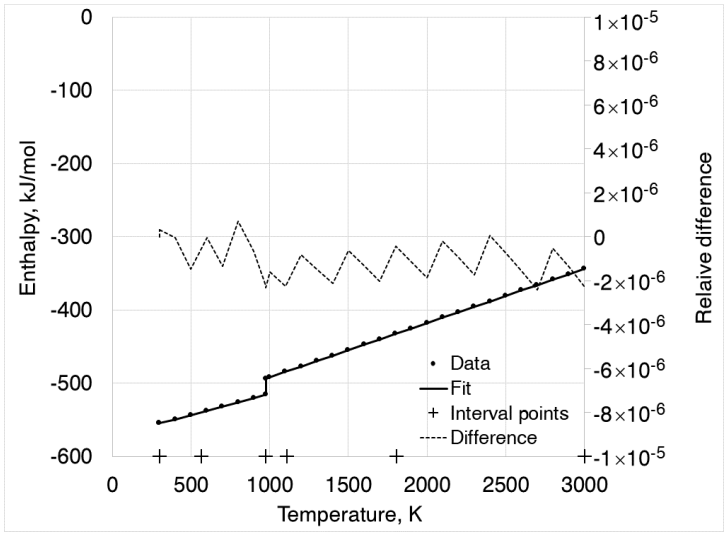


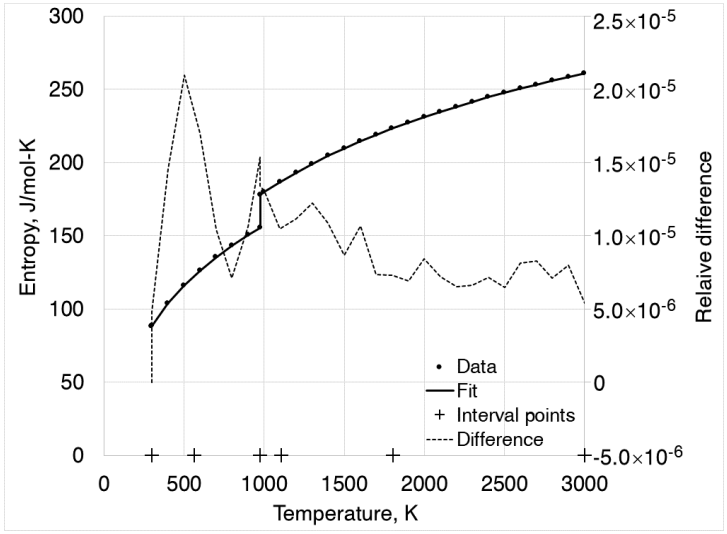


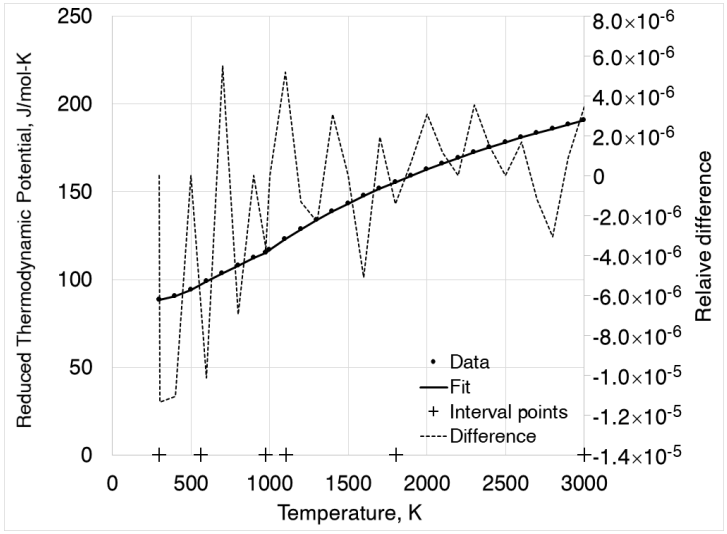
Cesium Fluoride (CsF) – Cs1F1(cr,l)

MW = 151.9038 g/mol – $\Delta_f H^o$(298.15 K) = -554.673 kJ/mol – S^o(298.15 K) = 88.282 J/mol-K – p^o = 0.1 MPa



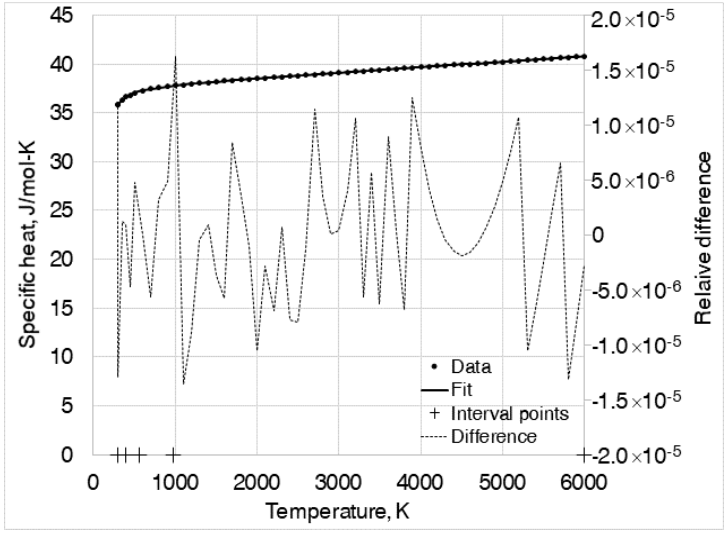


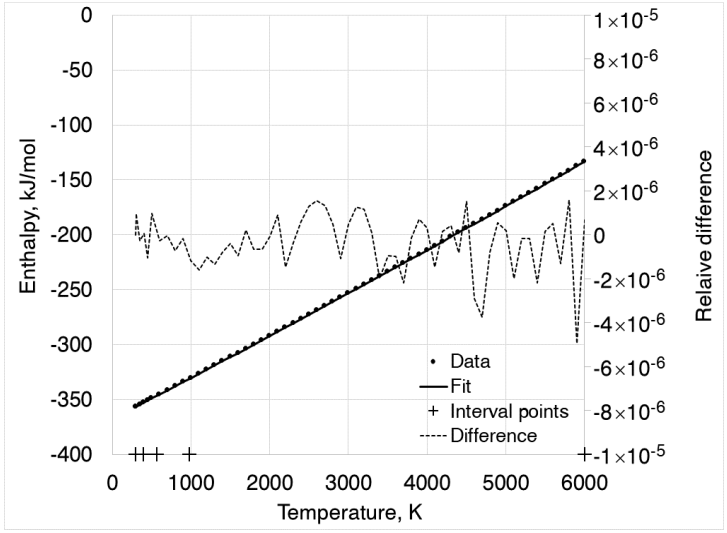


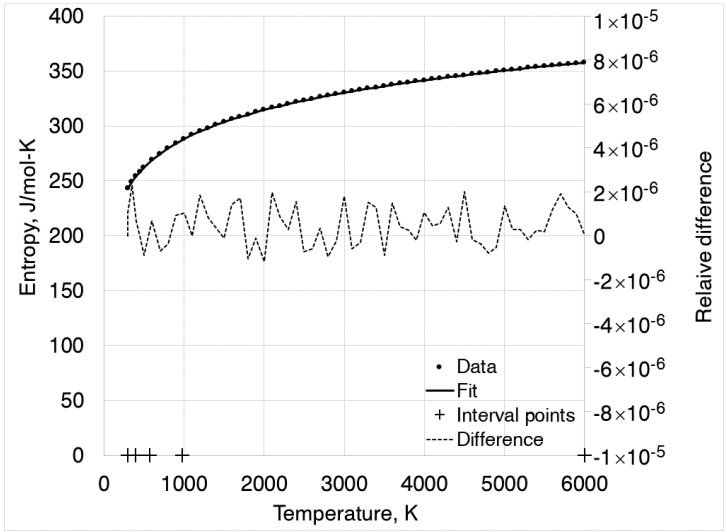


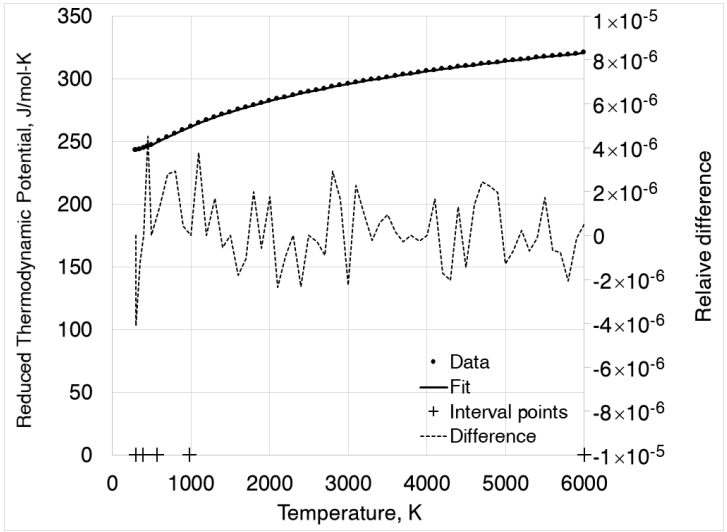
Cesium Fluoride (CsF) – Cs1F1(g)

MW = 151.9038 g/mol – $\Delta_f H^o$(298.15 K) = -356.477 kJ/mol – S^o(298.15 K) = 243.213 J/mol-K – p^o = 0.1 MPa



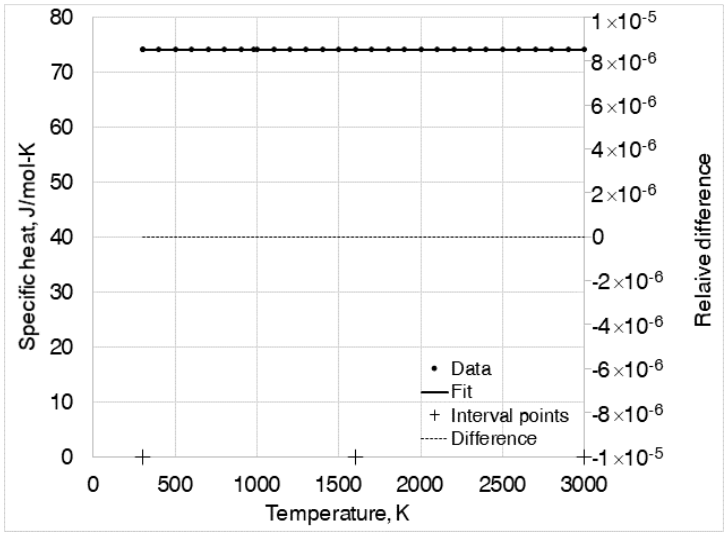


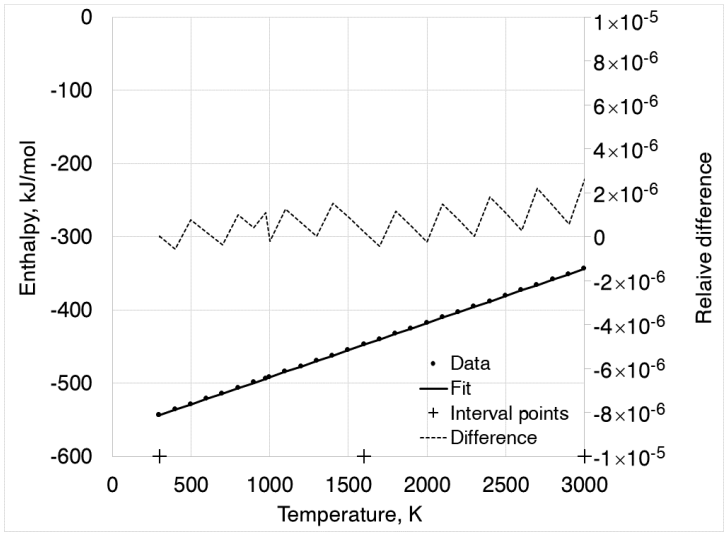


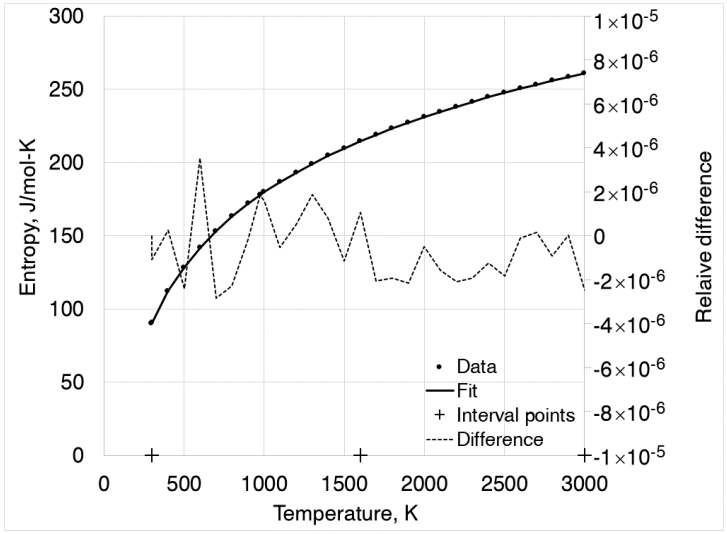


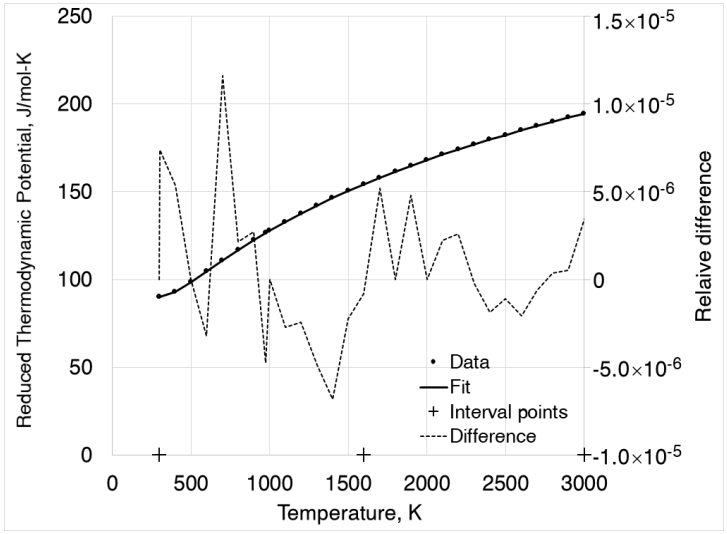
Cesium Fluoride (CsF) – Cs1F1(l)

MW = 151.9038 g/mol – $\Delta_f H^o$(298.15 K) = -543.849 kJ/mol – S^o(298.15 K) = 90.099 J/mol-K – p^o = 0.1 MPa



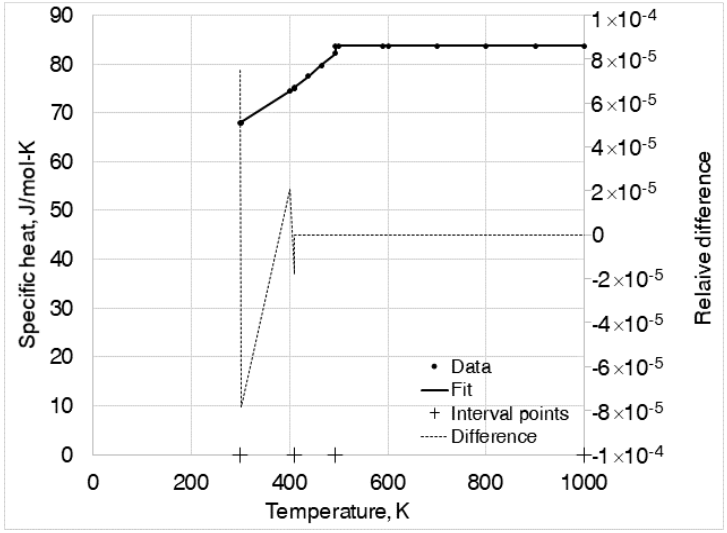


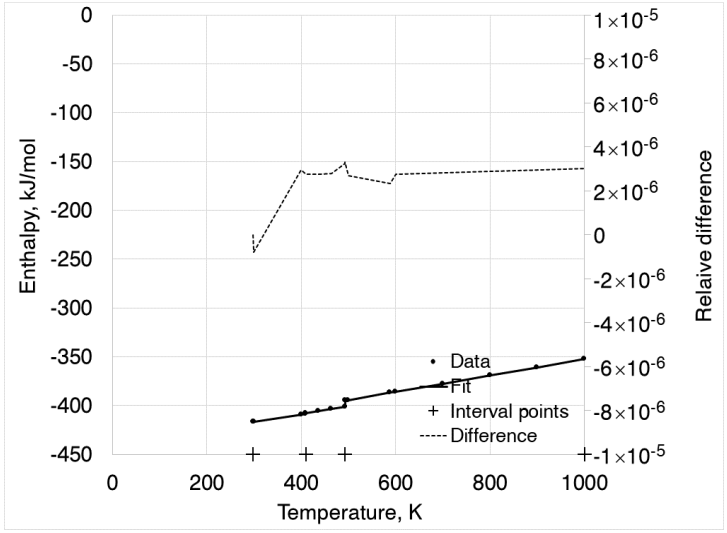


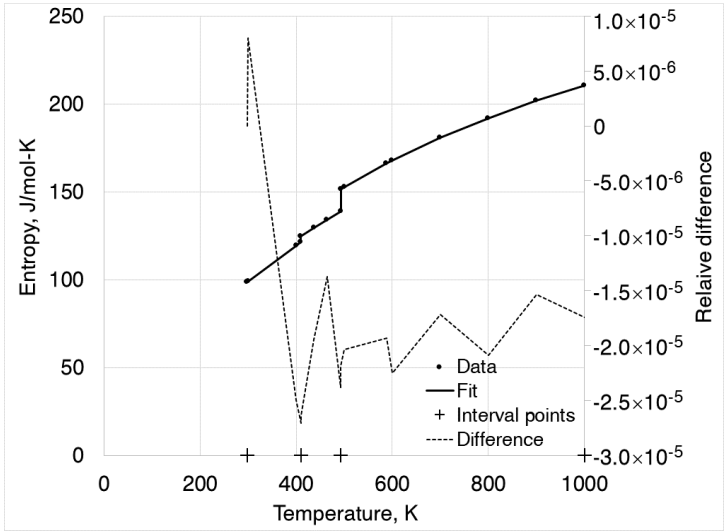


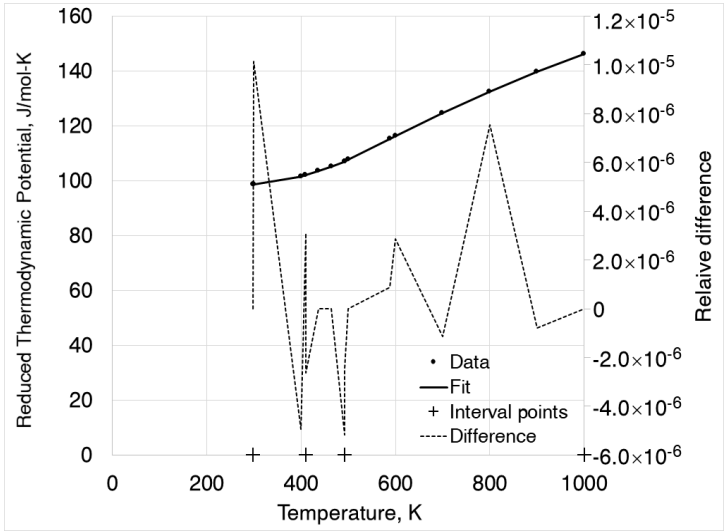
Cesium Hydroxide (CsOH) – Cs1H1O1(cr)

MW = 149.9127 g/mol – $\Delta_f H^o$(298.15 K) = -416.726 kJ/mol – S^o(298.15 K) = 98.742 J/mol-K – p^o = 0.1 MPa



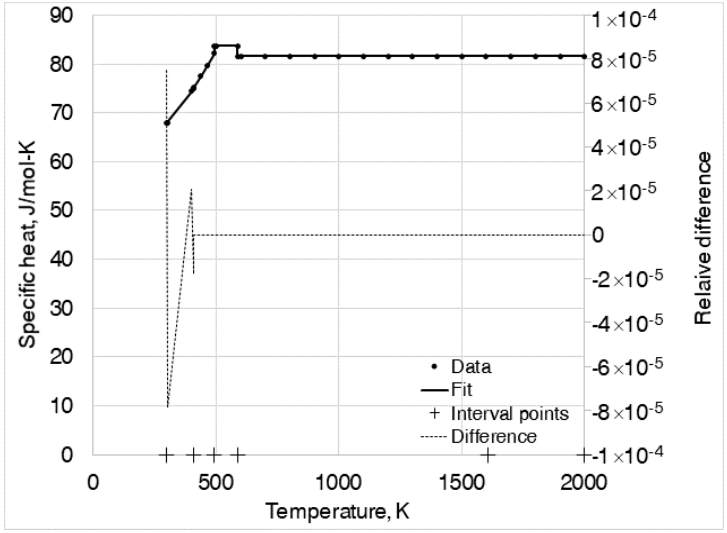


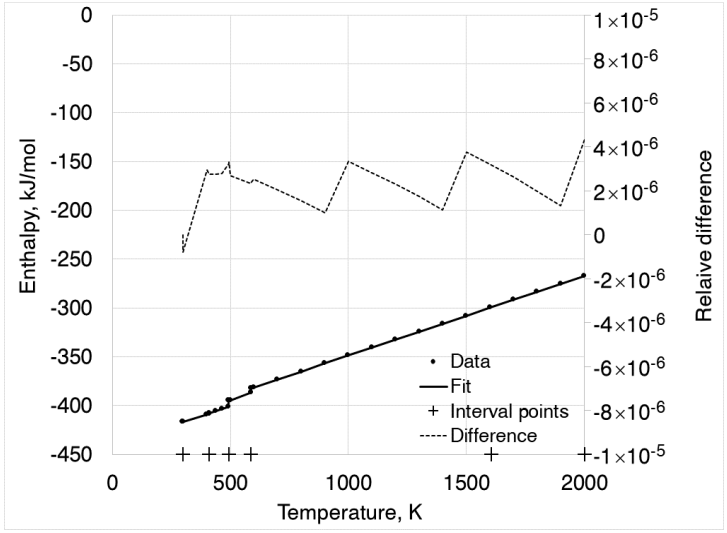


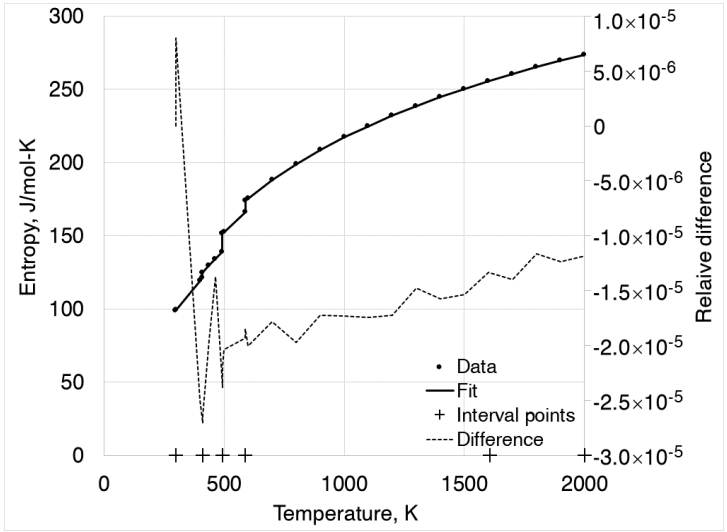


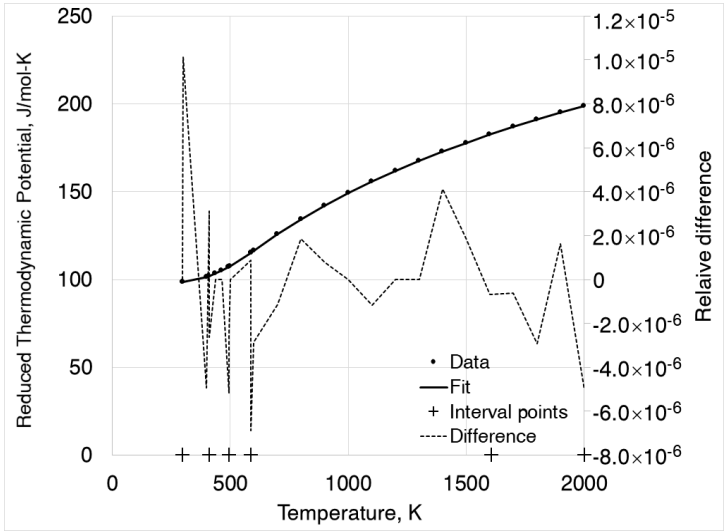
Cesium Hydroxide (CsOH) – Cs1H1O1(cr,l)

MW = 149.9127 g/mol – $\Delta_f H^o$(298.15 K) = -416.726 kJ/mol – S^o(298.15 K) = 98.742 J/mol-K – p^o = 0.1 MPa



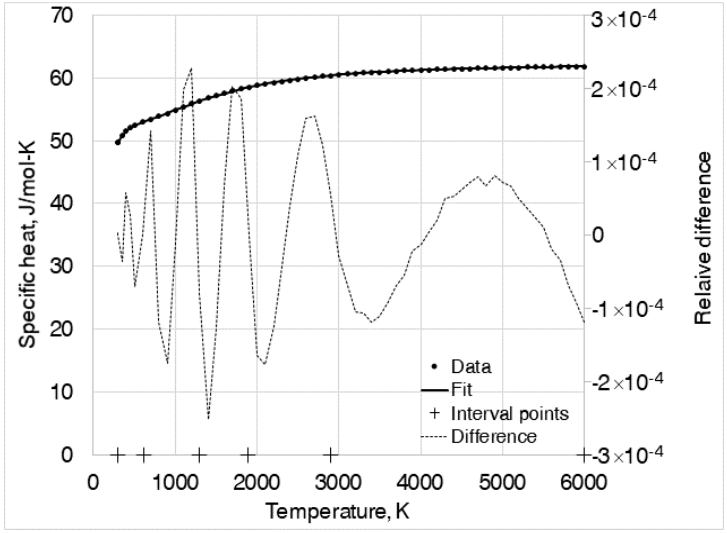


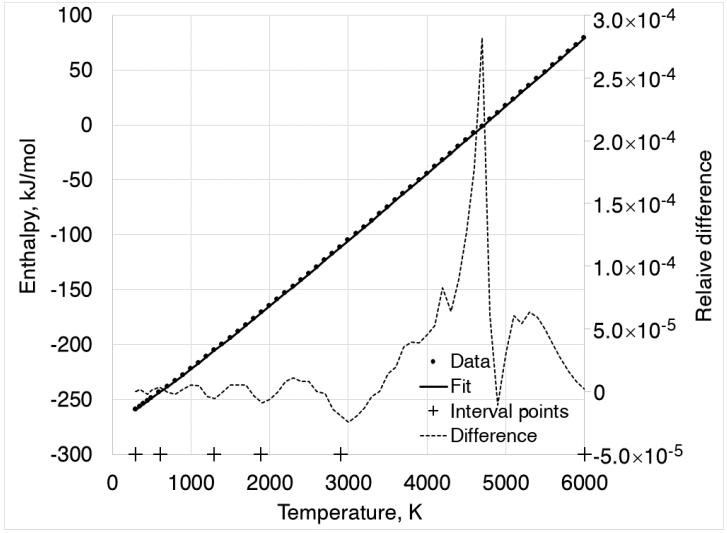


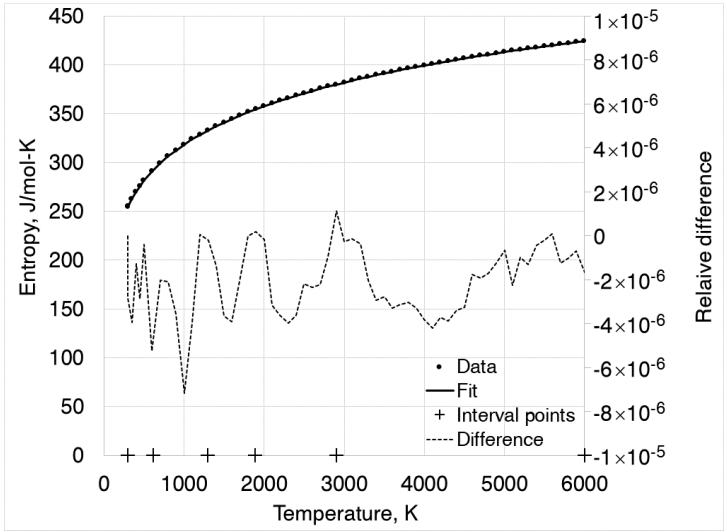


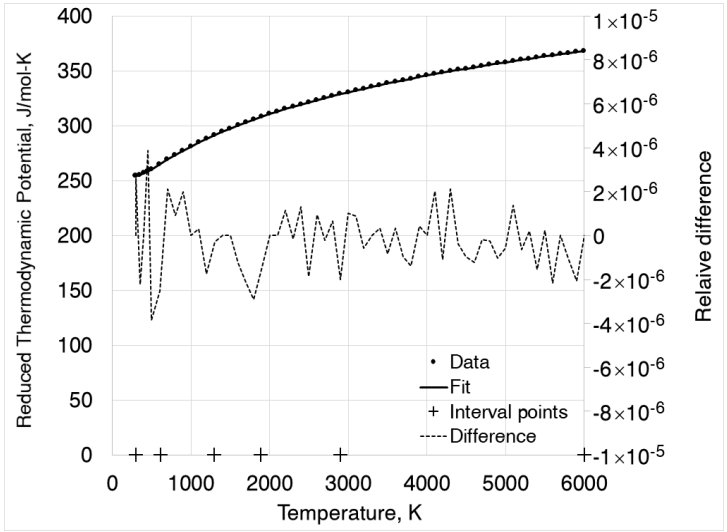
Cesium Hydroxide (CsOH) – Cs1H1O1(g)

MW = 149.9127 g/mol – $\Delta_f H^o$(298.15 K) = -259.408 kJ/mol – S^o(298.15 K) = 254.778 J/mol-K – p^o = 0.1 MPa



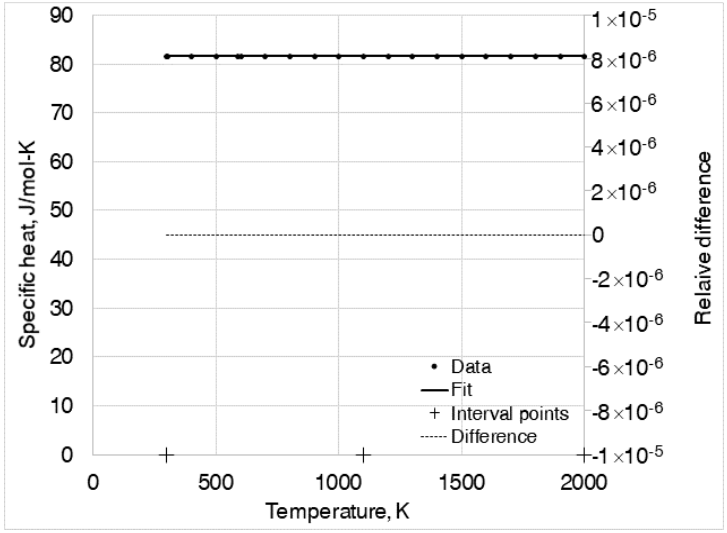


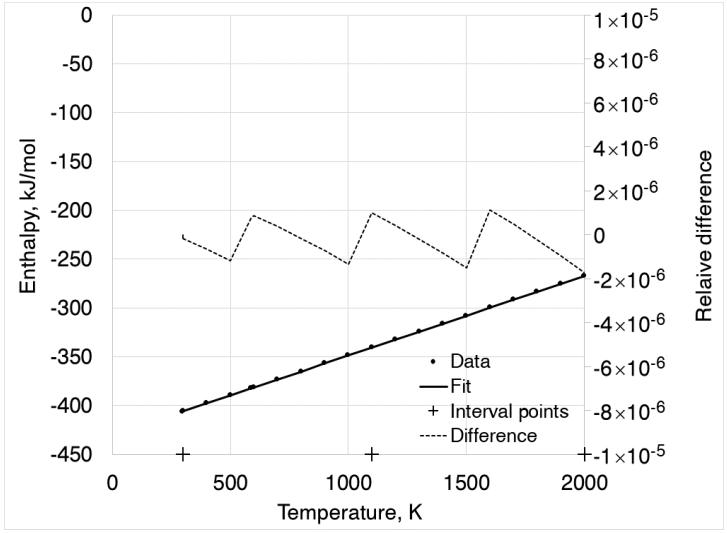


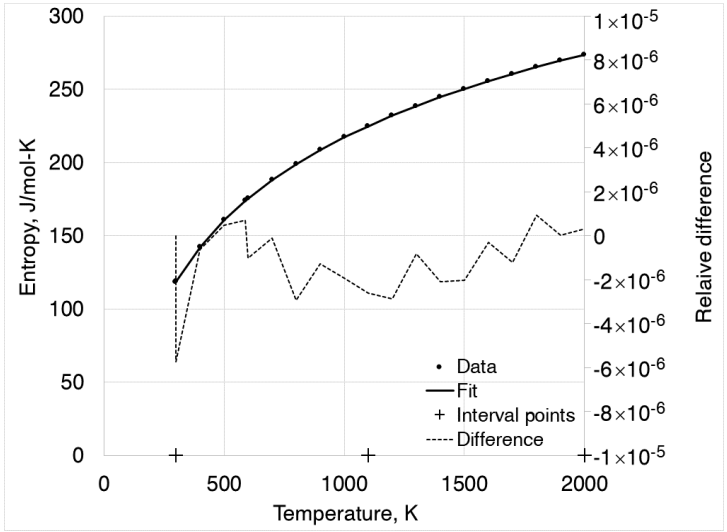


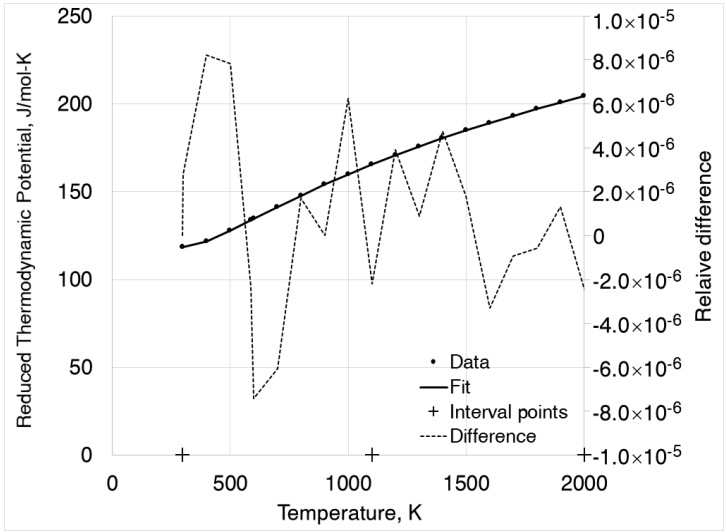
Cesium Hydroxide (CsOH) – Cs1H1O1(l)

MW = 149.9127 g/mol – $\Delta_f H^o$(298.15 K) = -405.97 kJ/mol – S^o(298.15 K) = 118.513 J/mol-K – p^o = 0.1 MPa



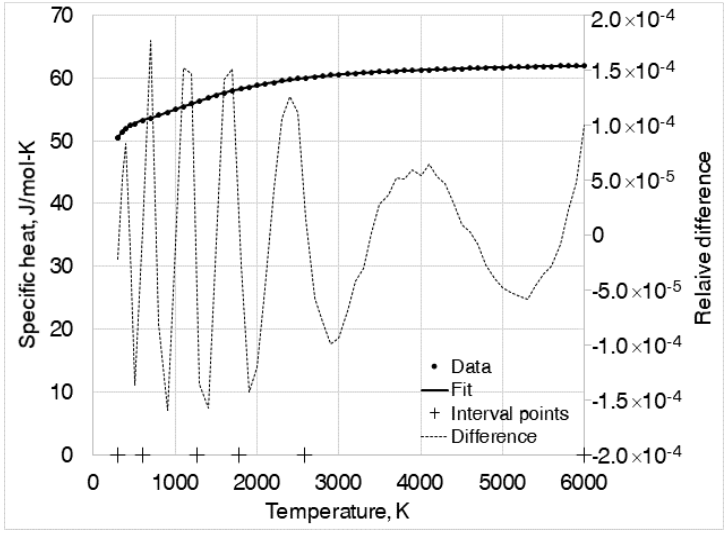


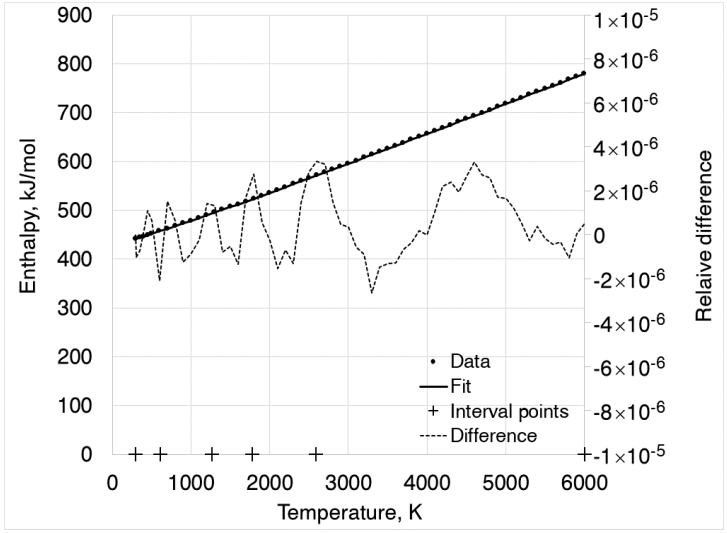


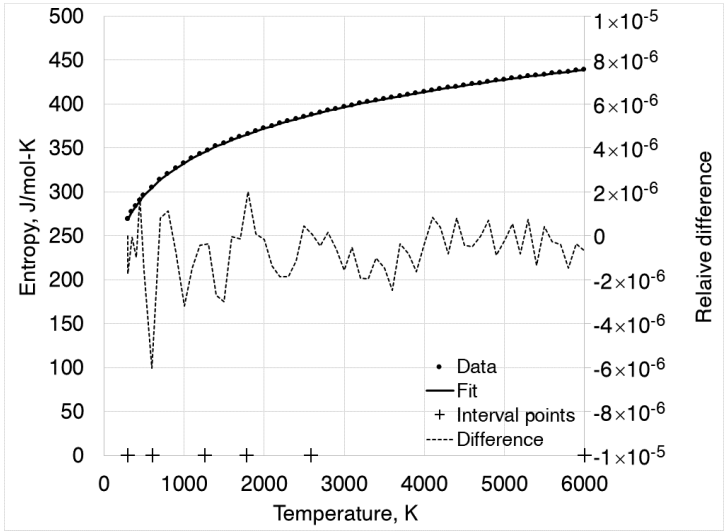


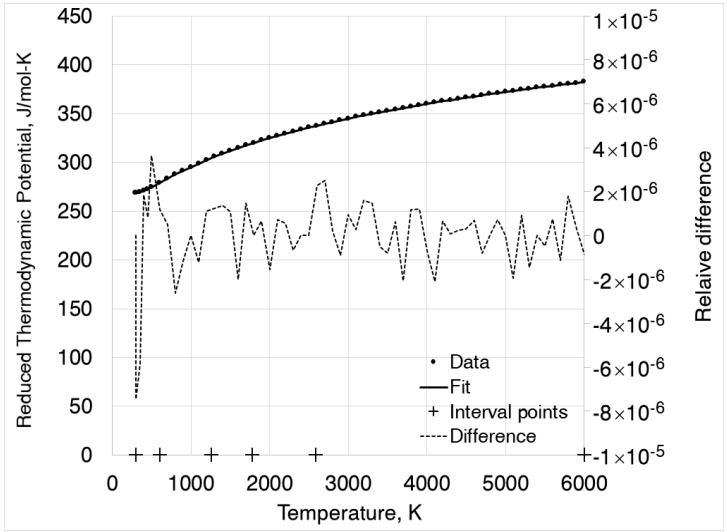
Cesium Hydroxide, Ion ($CsOH^+$) – Cs1H1O1+(g)

MW = 149.9122 g/mol – $\Delta_f H^o$(298.15 K) = 441.379 kJ/mol – S^o(298.15 K) = 268.686 J/mol-K – p^o = 0.1 MPa



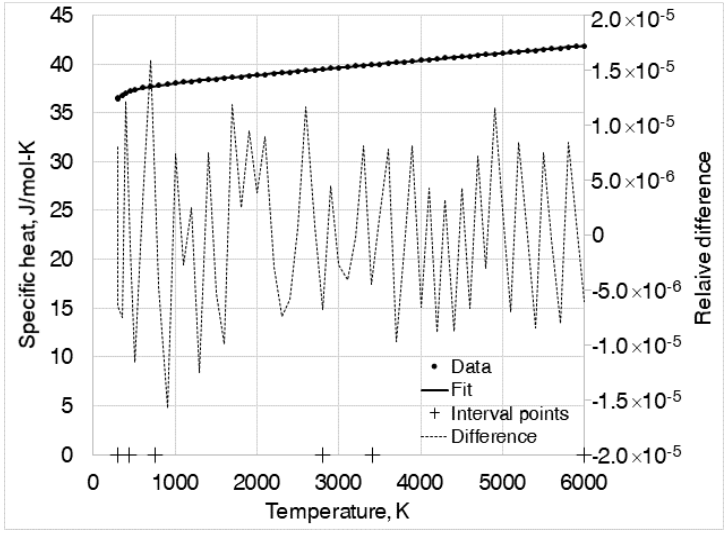


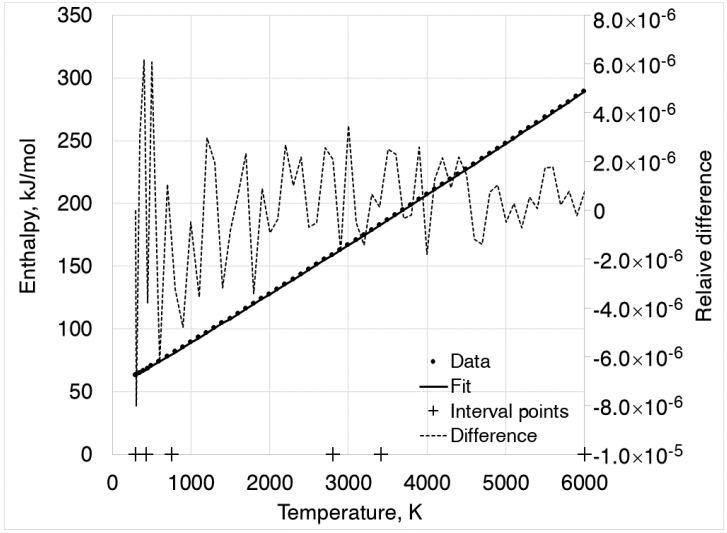


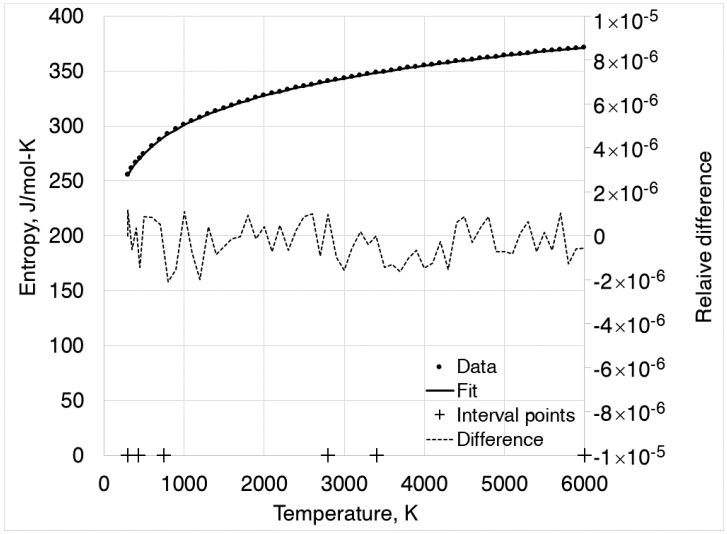


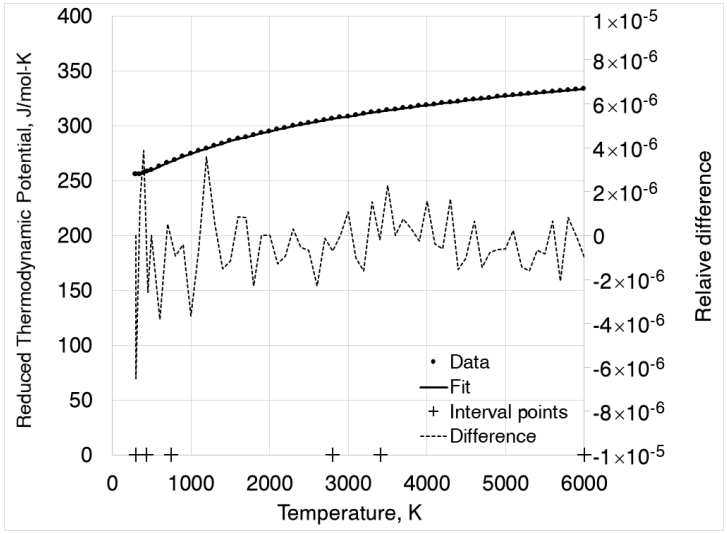
Cesium Oxide (CsO) – Cs1O1(g)

MW = 148.9048 g/mol – $\Delta_f H^o$(298.15 K) = 62.76 kJ/mol – S^o(298.15 K) = 255.542 J/mol-K – p^o = 0.1 MPa



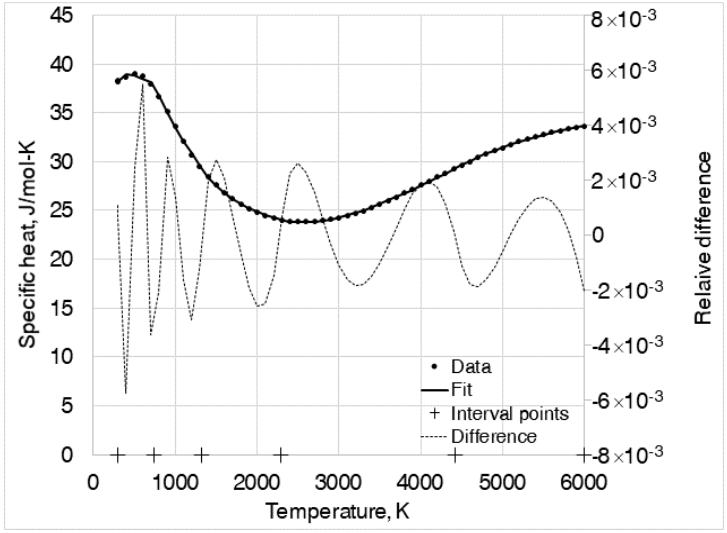


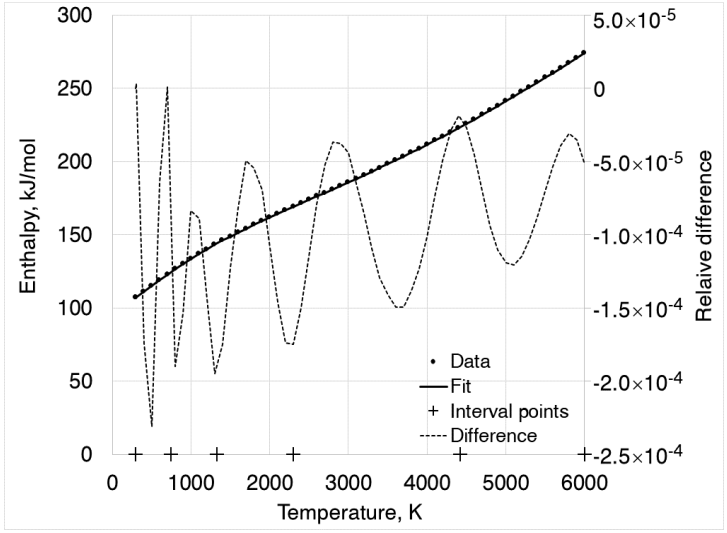


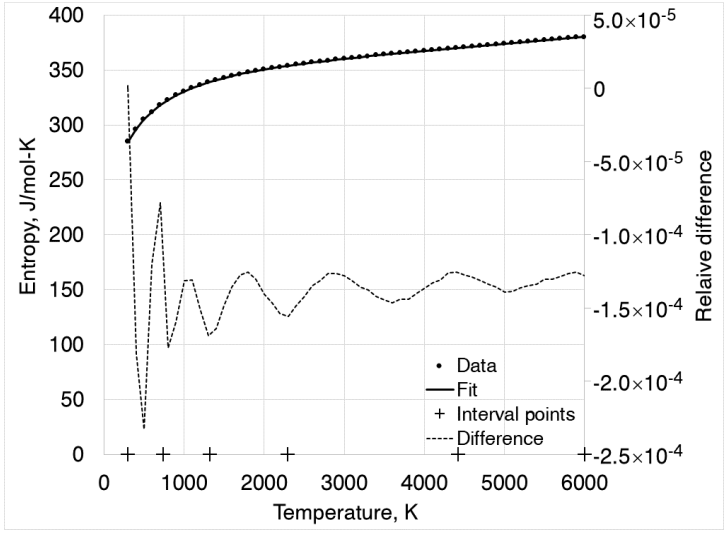


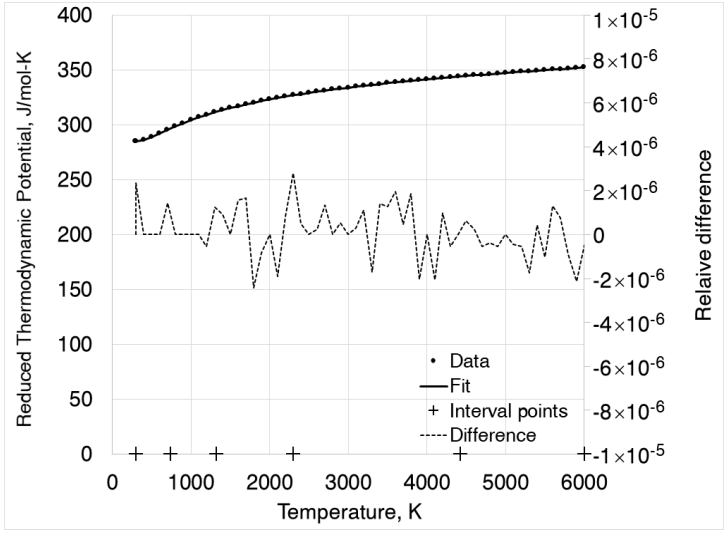
Cesium (Cs_2) – Cs2(g)

MW = 265.8108 g/mol – $\Delta_f H^o$(298.15 K) = 107.366 kJ/mol – S^o(298.15 K) = 284.675 J/mol-K – p^o = 0.1 MPa



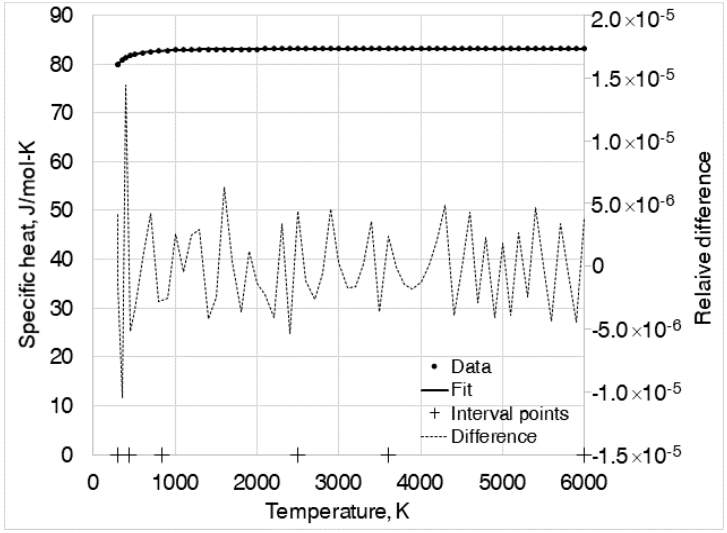


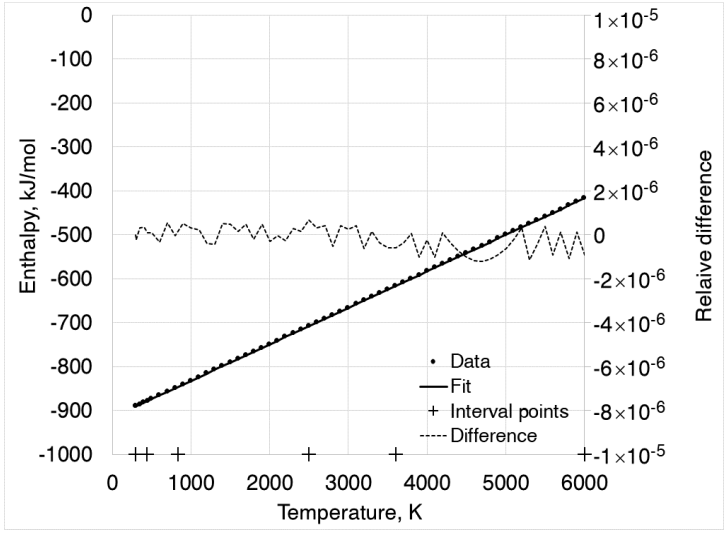


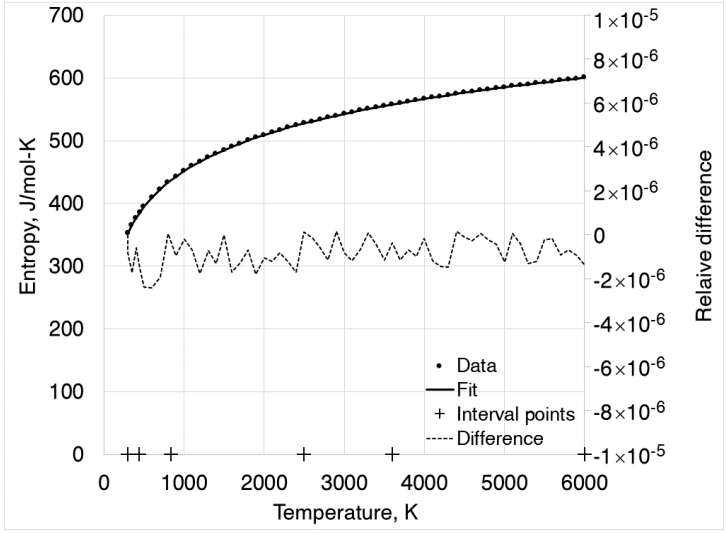


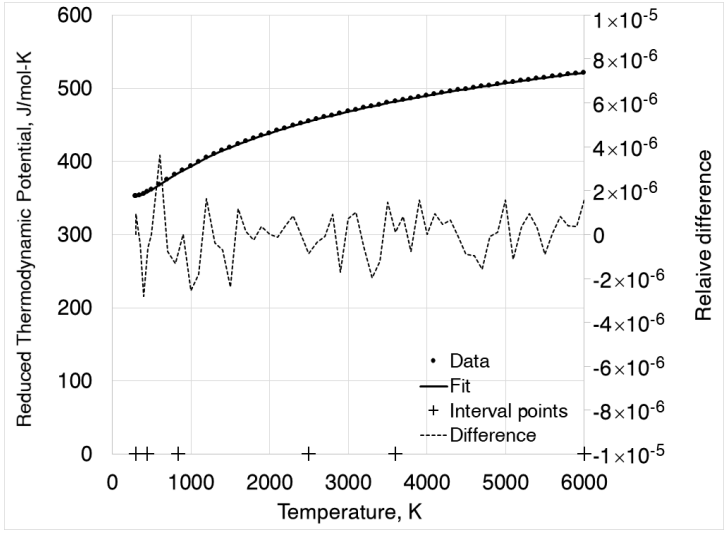
Cesium Fluoride ((CsF)$_2$) – Cs2F2(g)

MW = 303.8076 g/mol – $\Delta_f H^o$(298.15 K) = -890.104 kJ/mol – S^o(298.15 K) = 352.319 J/mol-K – p^o = 0.1 MPa



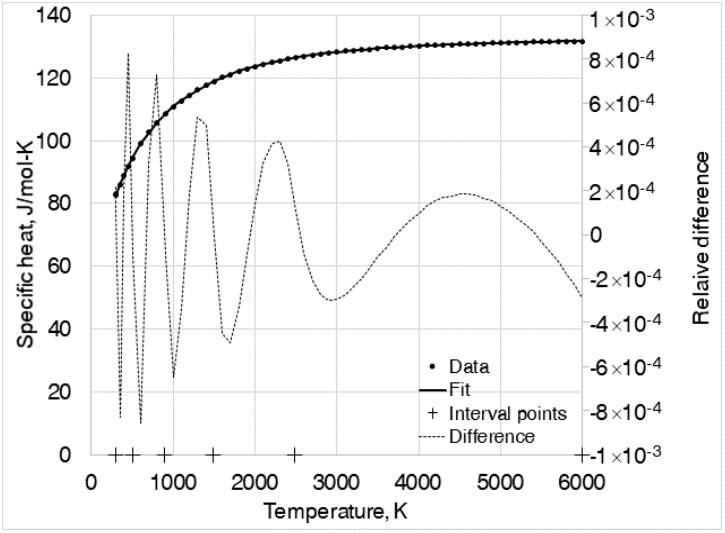


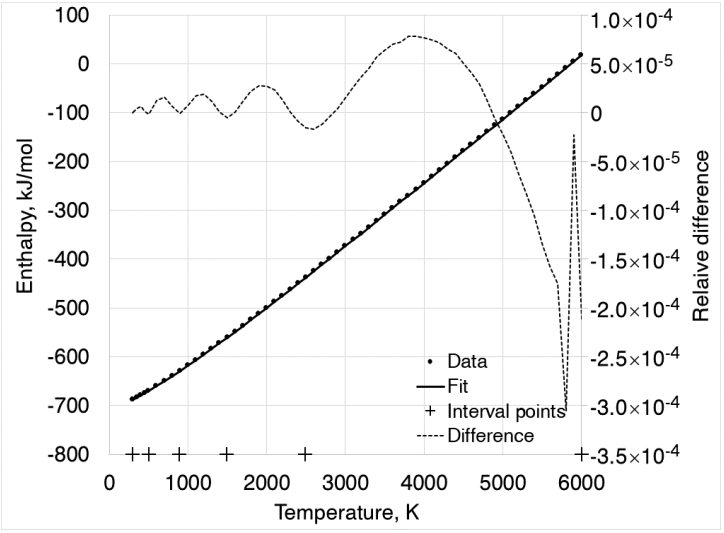


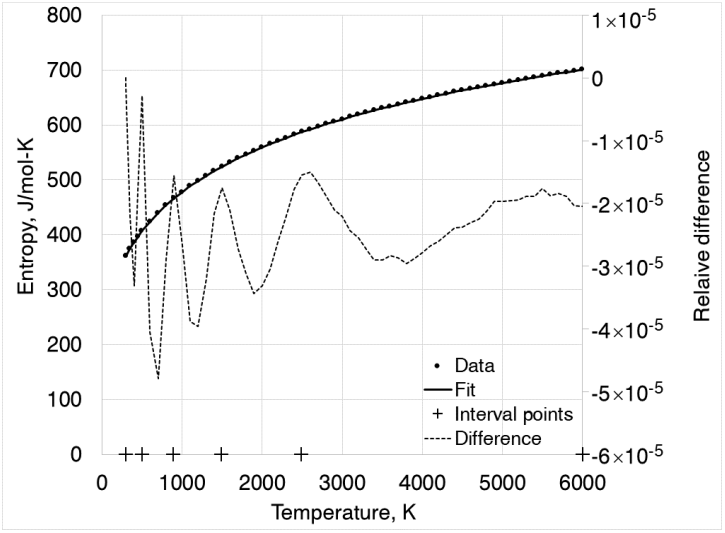


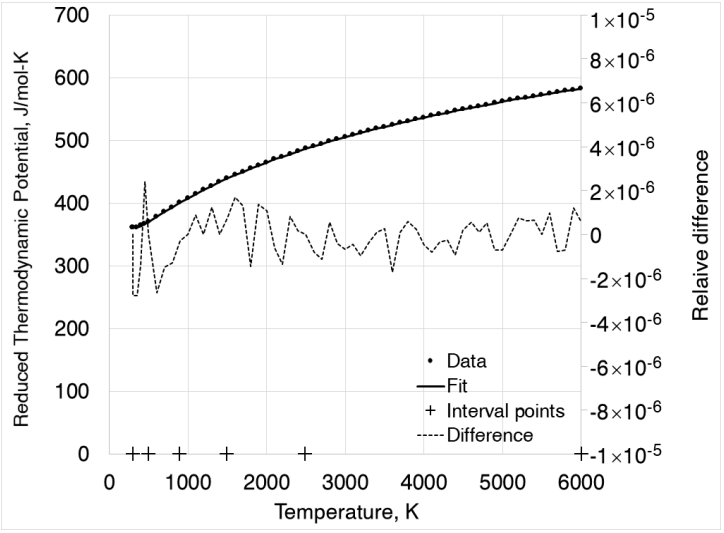
Cesium Hydroxide ((CsOH)$_2$) – Cs2H2O2(g)

MW = 299.8255 g/mol – $\Delta_f H^o$(298.15 K) = -687.85 kJ/mol – S^o(298.15 K) = 360.748 J/mol-K – p^o = 0.1 MPa



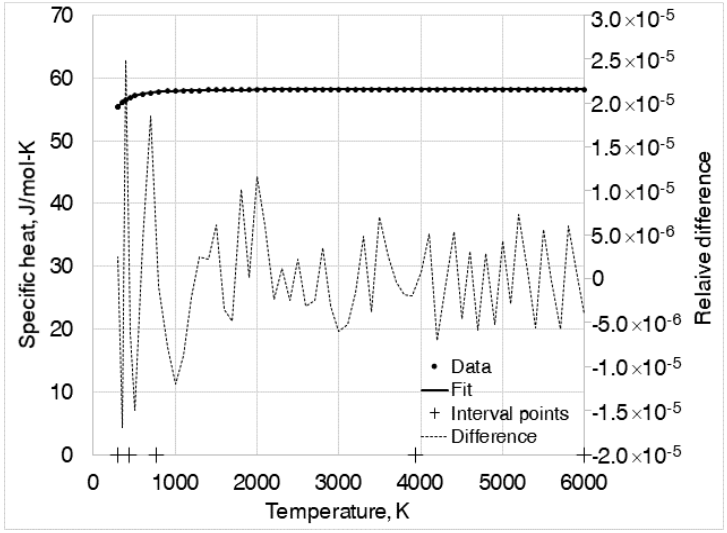


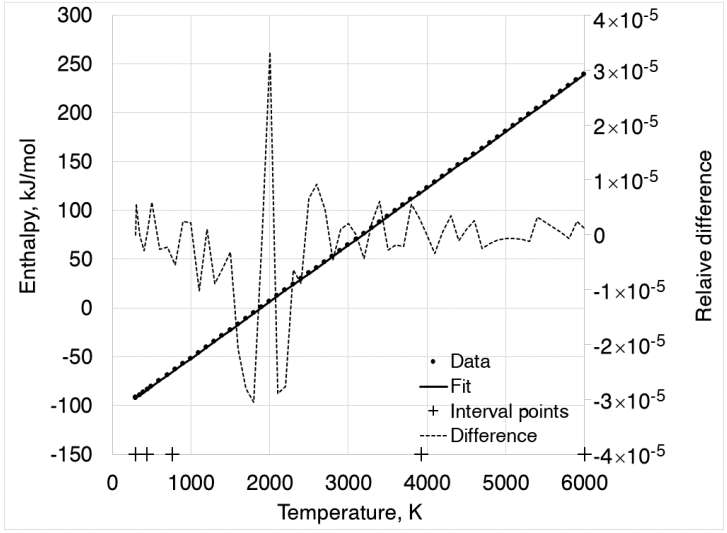


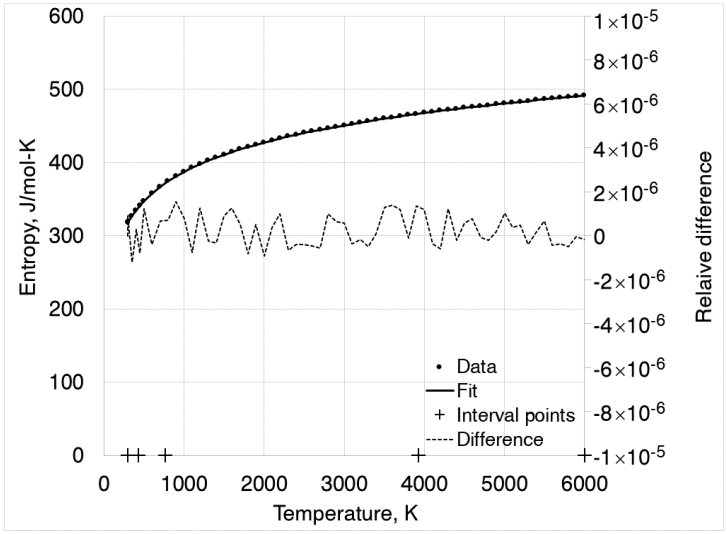


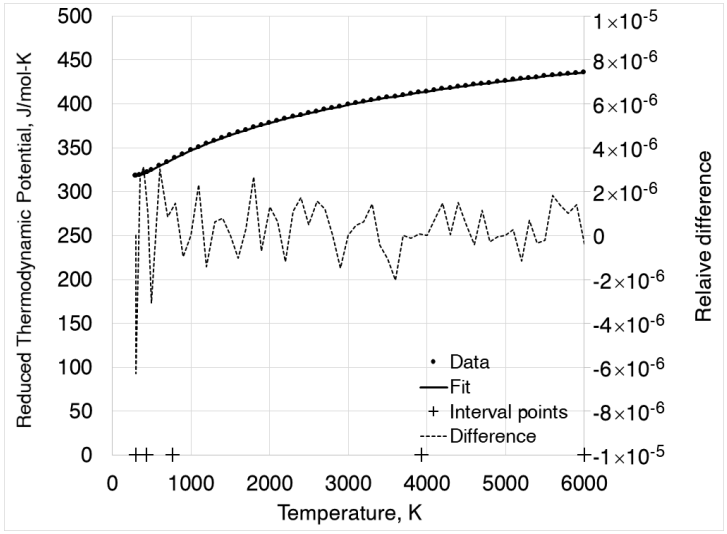
Cesium Oxide (Cs_2O) – Cs2O1(g)

MW = 281.8102 g/mol – $\Delta_f H^o$(298.15 K) = -92.048 kJ/mol – S^o(298.15 K) = 318.088 J/mol-K – p^o = 0.1 MPa



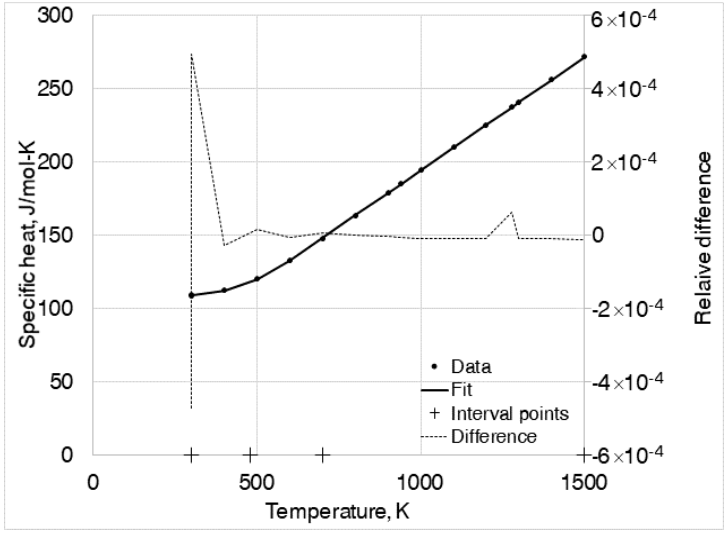


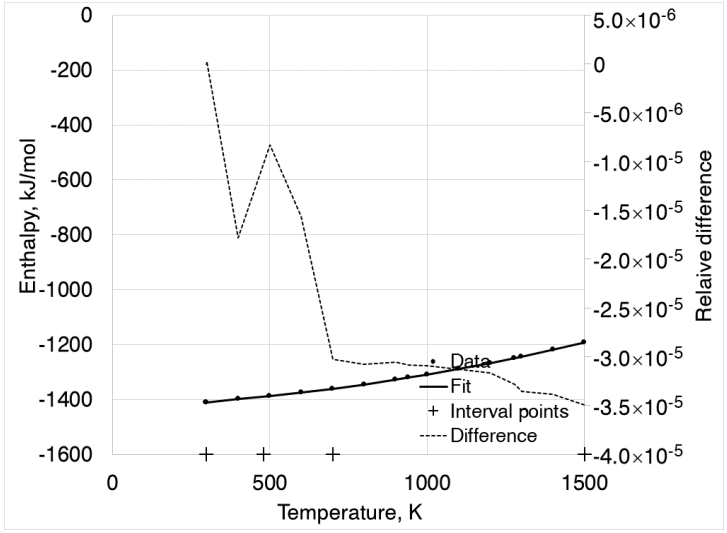


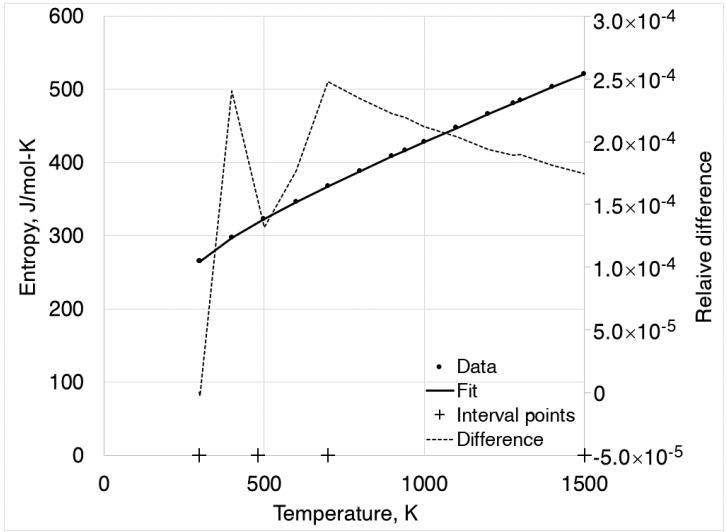


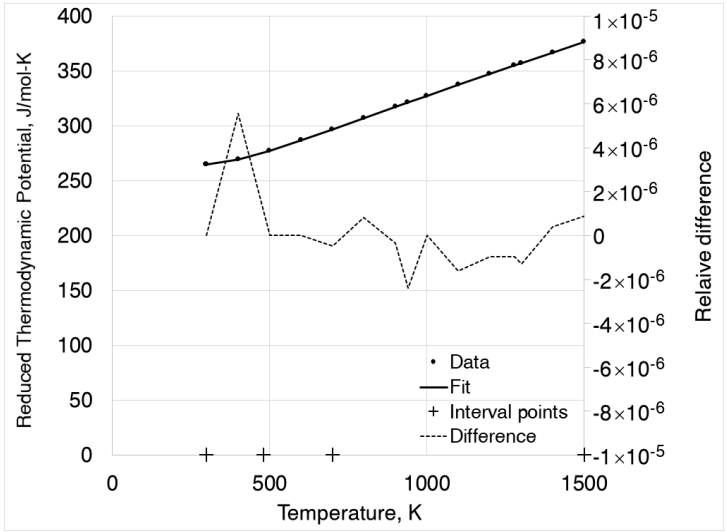
Cesium Sulfate, I (Cs_2SO_4) – Cs2O4S1(cr)_01

MW = 361.87 g/mol – $\Delta_f H^o$(298.15 K) = -1410.152 kJ/mol – S^o(298.15 K) = 264.834 J/mol-K – p^o = 0.1 MPa



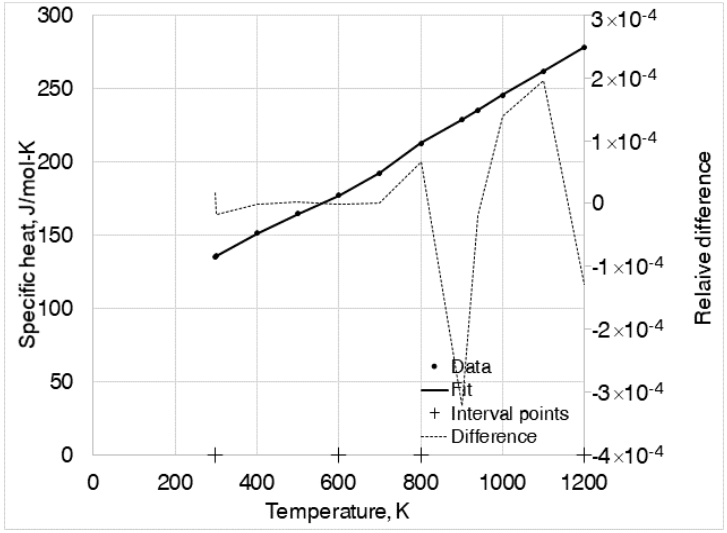


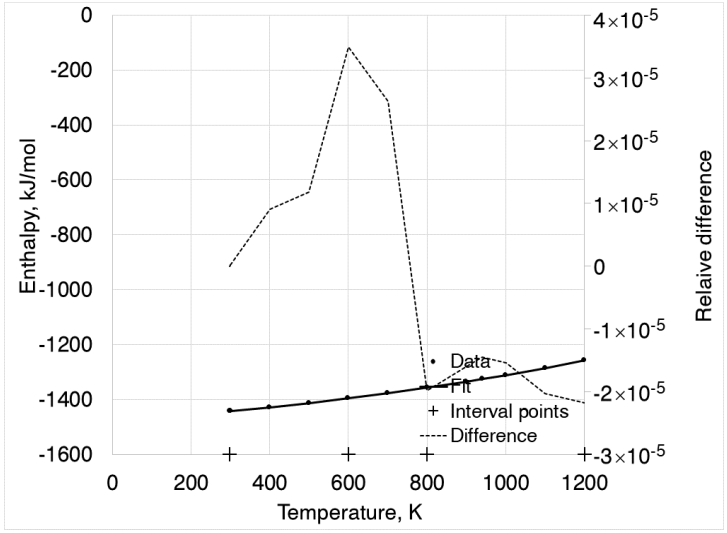


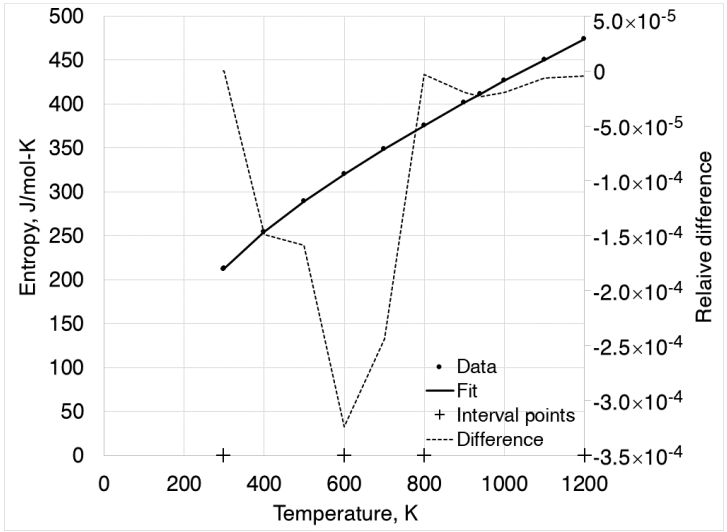


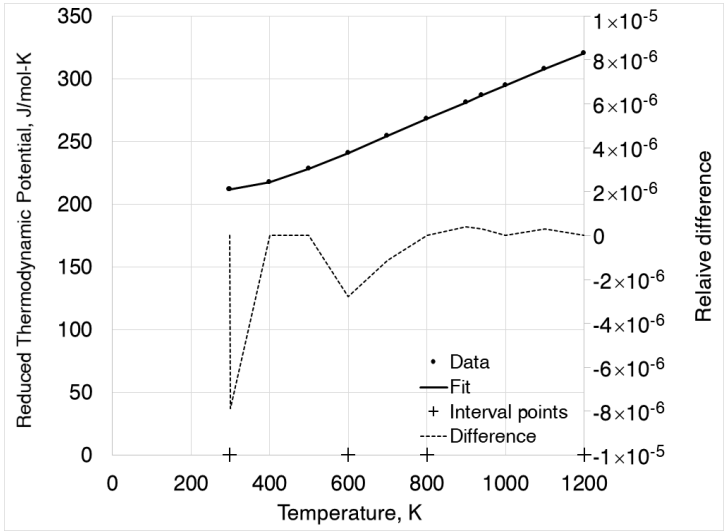
Cesium Sulfate, II (Cs_2SO_4) – Cs2O4S1(cr)_02

MW = 361.87 g/mol – $\Delta_f H^o$(298.15 K) = -1442.643 kJ/mol – S^o(298.15 K) = 211.878 J/mol-K – p^o = 0.1 MPa



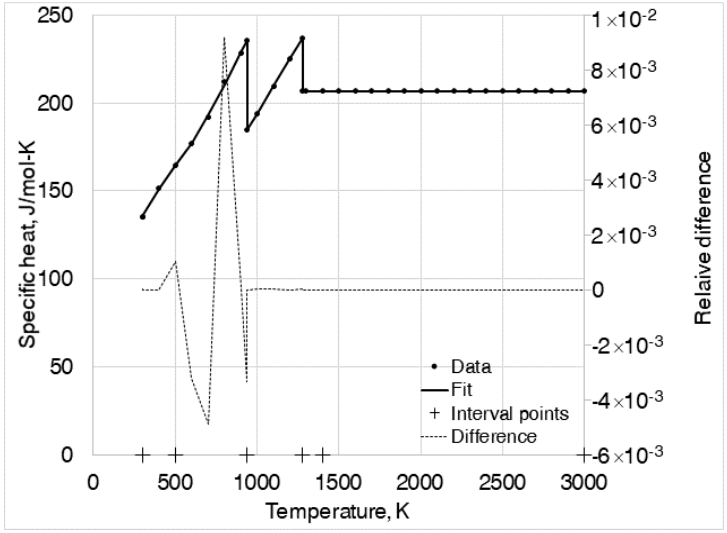


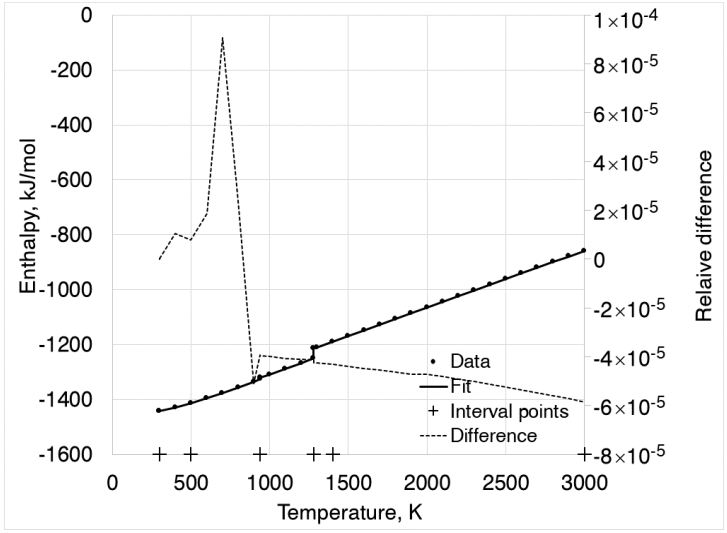


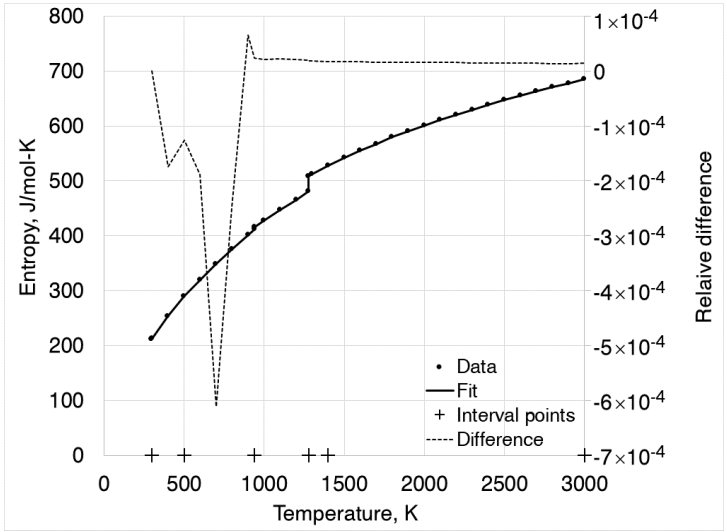


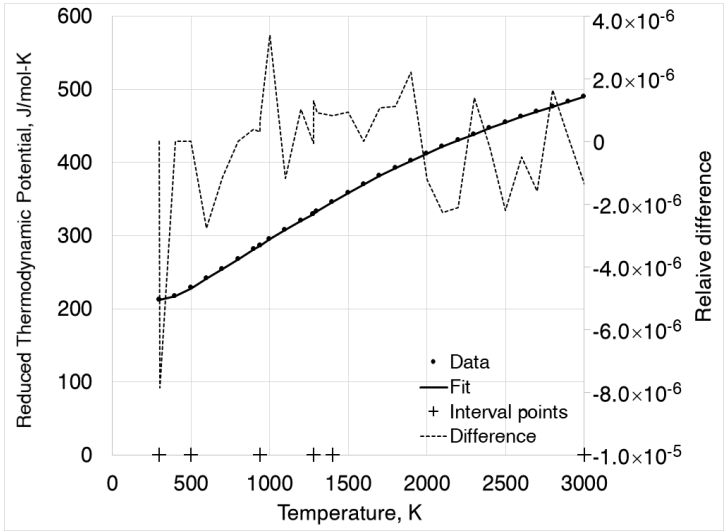
Cesium Sulfate (Cs_2SO_4) – Cs2O4S1(cr,l)

MW = 361.87 g/mol – $\Delta_f H^o$(298.15 K) = -1442.643 kJ/mol – S^o(298.15 K) = 211.878 J/mol-K – p^o = 0.1 MPa



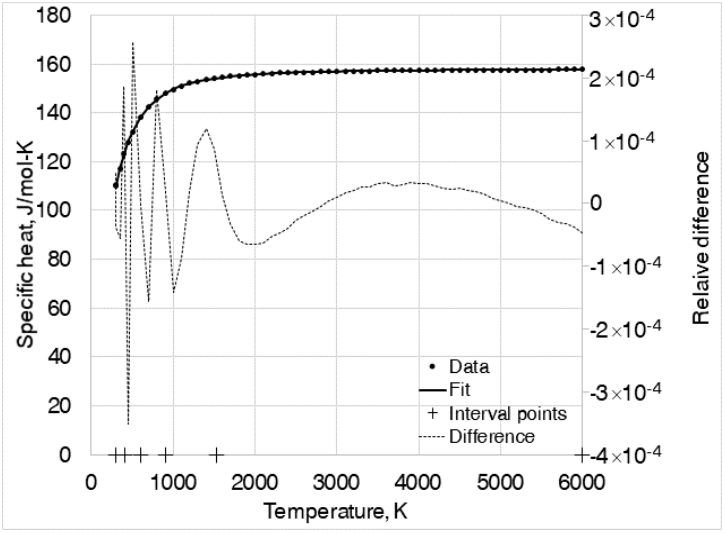


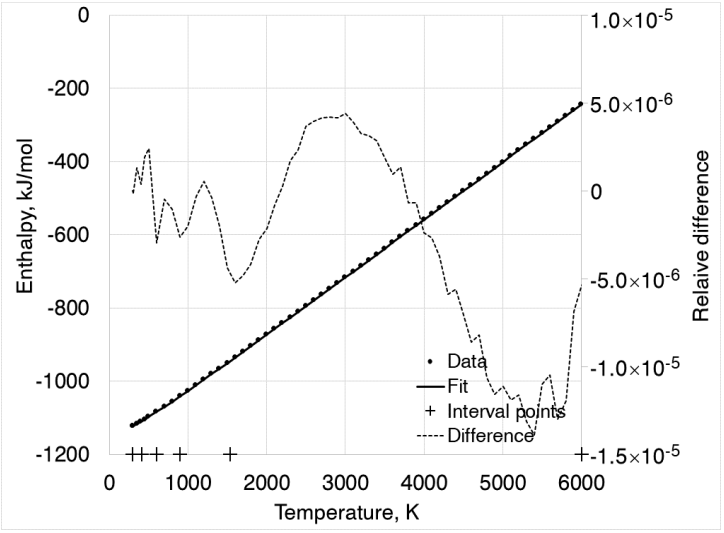


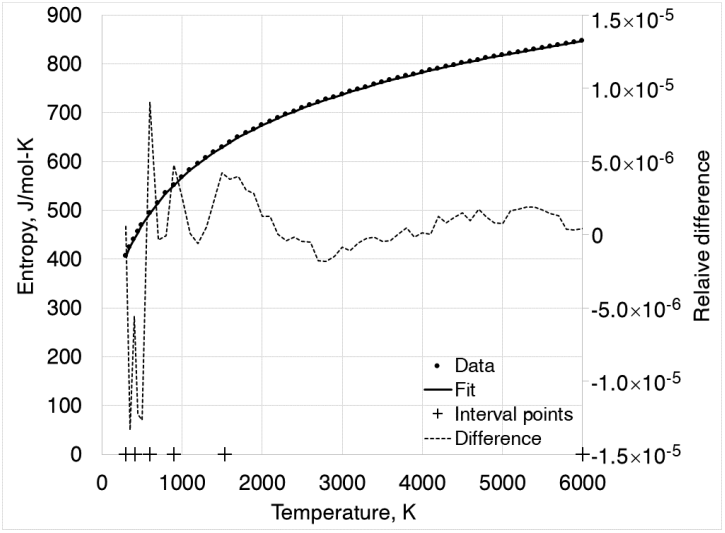


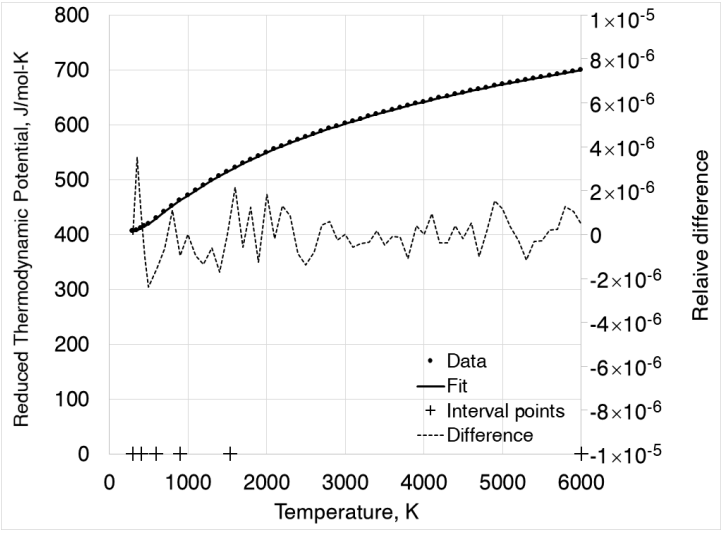
Cesium Sulfate (Cs_2SO_4) – Cs2O4S1(g)

MW = 361.87 g/mol – $\Delta_f H^o$(298.15 K) = -1122.567 kJ/mol – S^o(298.15 K) = 406.484 J/mol-K – p^o = 0.1 MPa



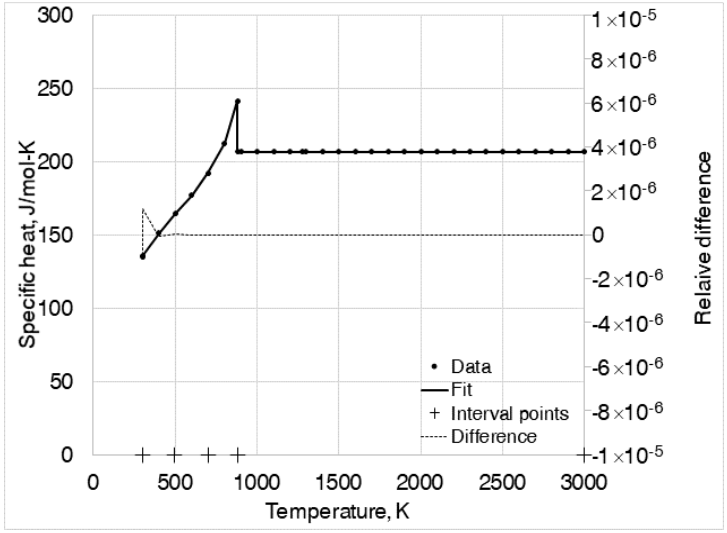


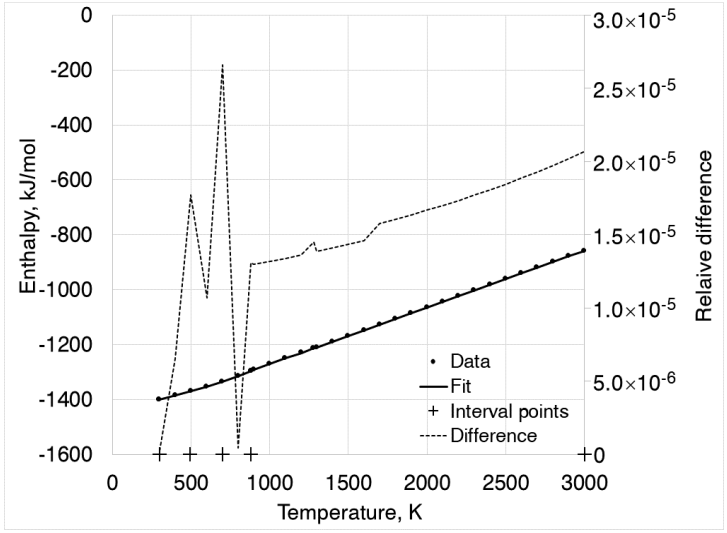


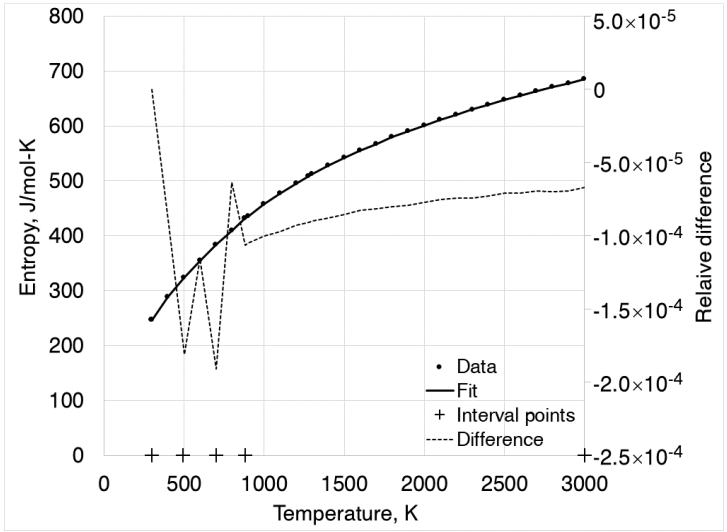


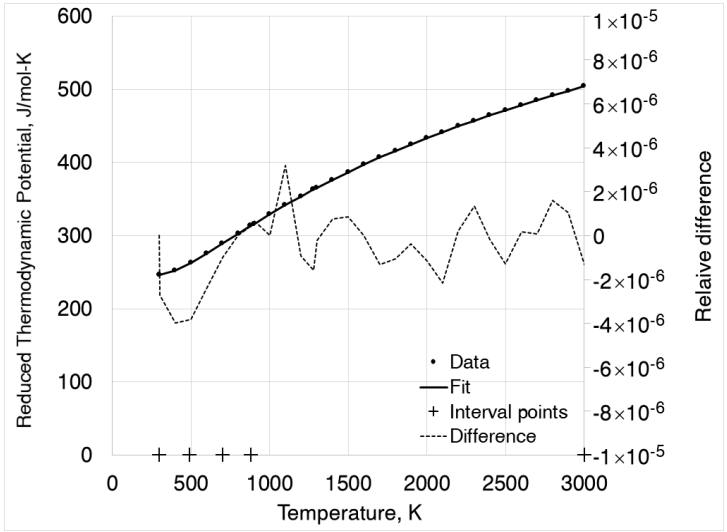
Cesium Sulfate (Cs_2SO_4) – Cs2O4S1(l)

MW = 361.87 g/mol – $\Delta_f H^o$(298.15 K) = -1400.139 kJ/mol – S^o(298.15 K) = 246.449 J/mol-K – p^o = 0.1 MPa



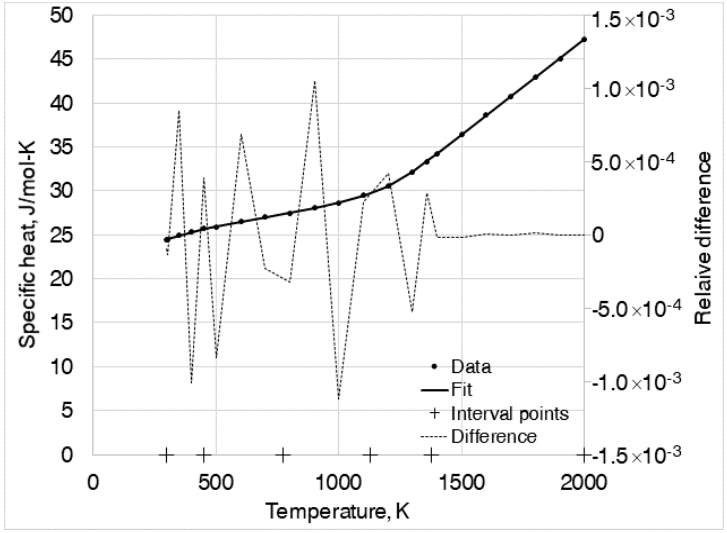


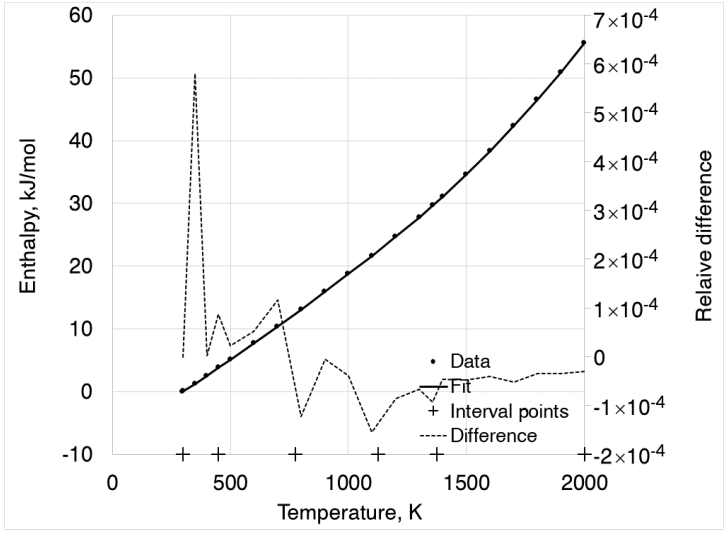


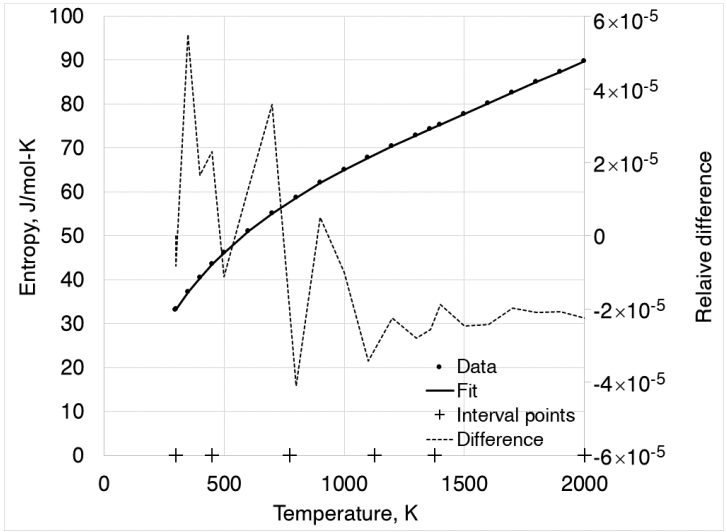


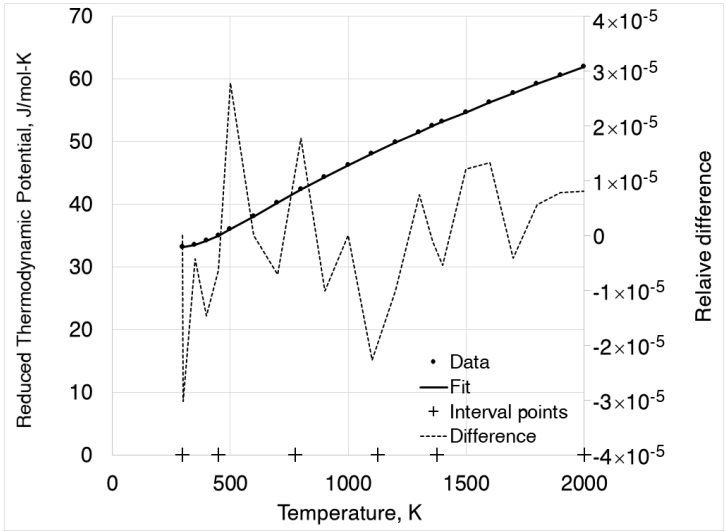
Copper (Cu) – Cu1(cr)

MW = 63.546 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 33.164 J/mol-K – p^o = 0.1 MPa



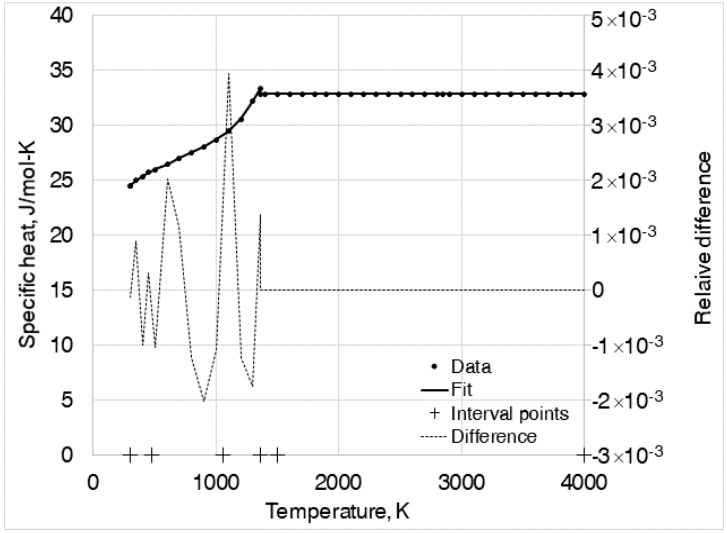


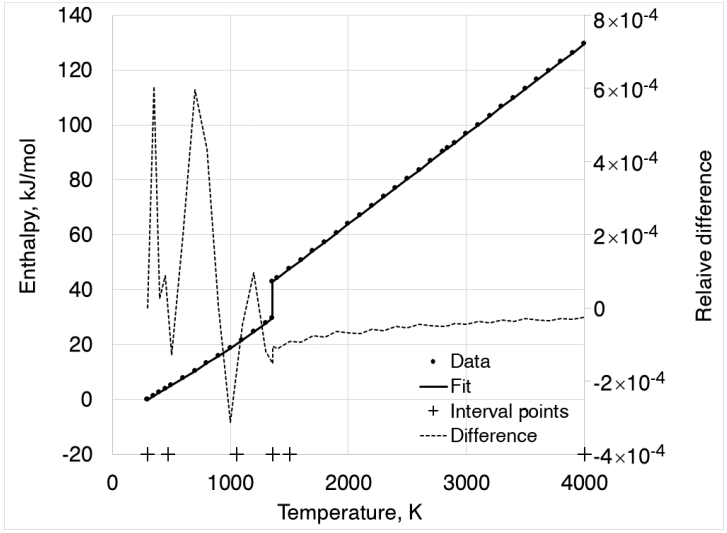


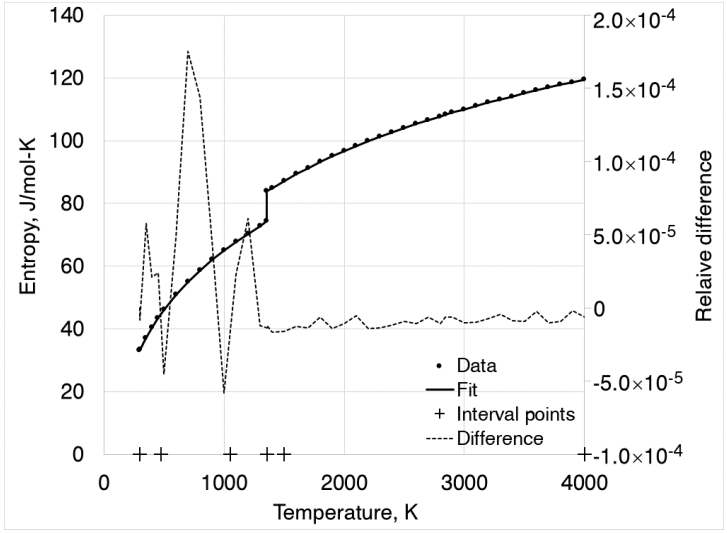


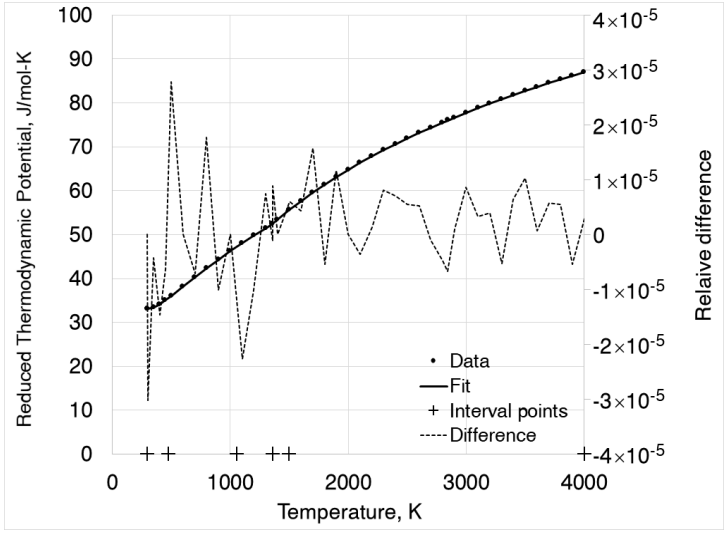
Copper (Cu) – Cu1(cr,l)

MW = 63.546 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 33.164 J/mol-K – p^o = 0.1 MPa



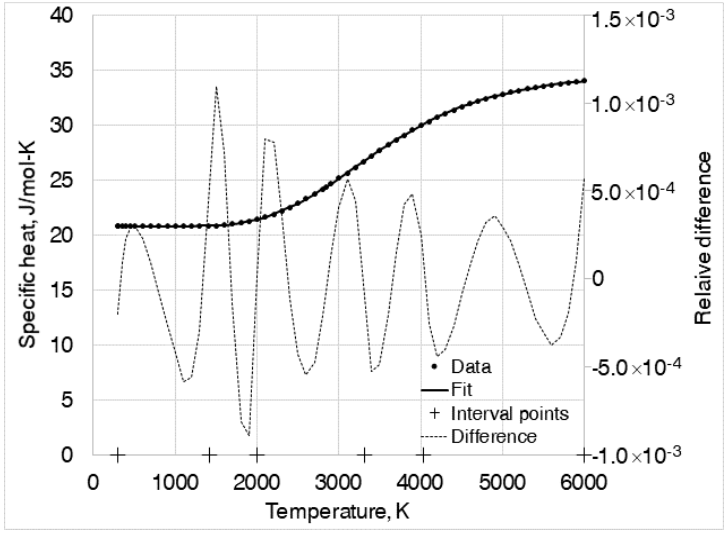


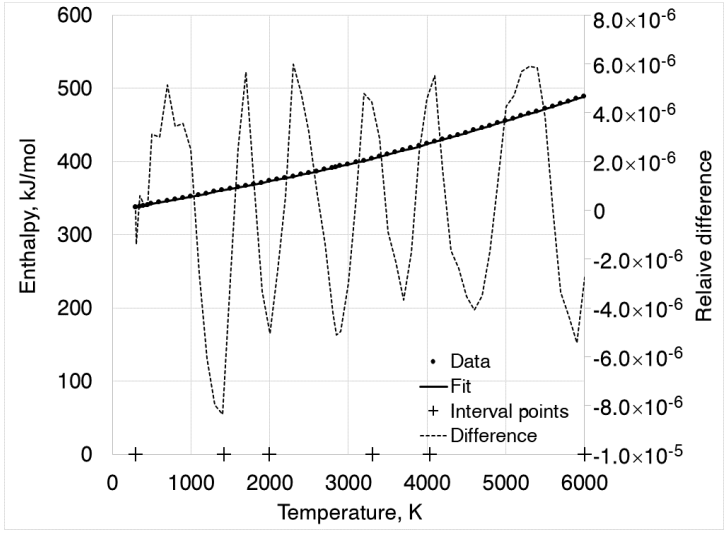


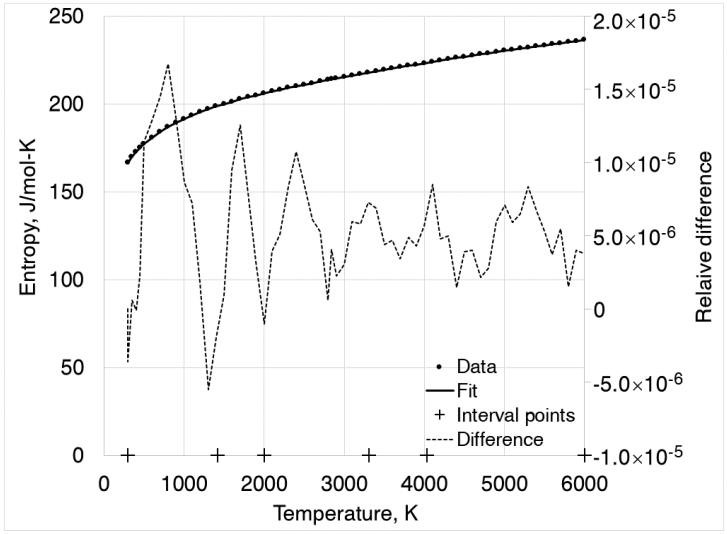


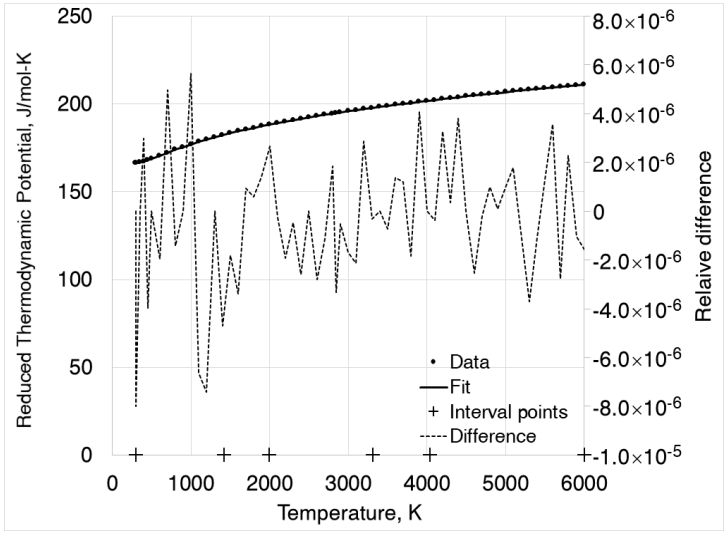
Copper (Cu) – Cu1(g)

MW = 63.546 g/mol – $\Delta_f H^o$(298.15 K) = 337.6 kJ/mol – S^o(298.15 K) = 166.397 J/mol-K – p^o = 0.1 MPa



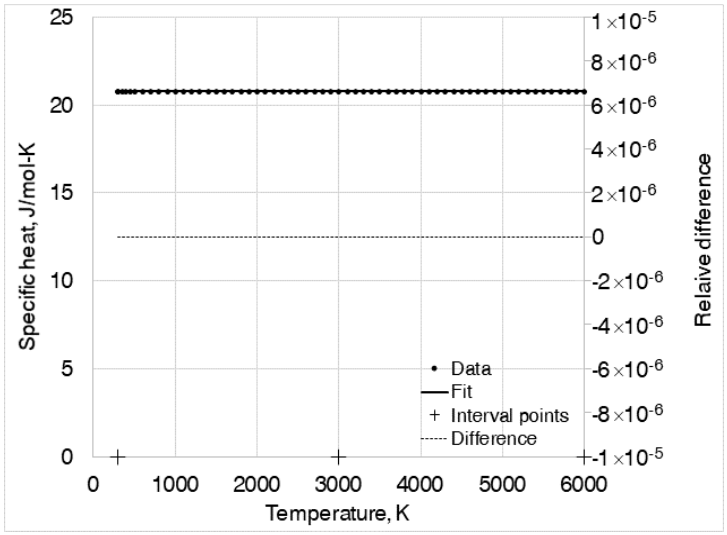


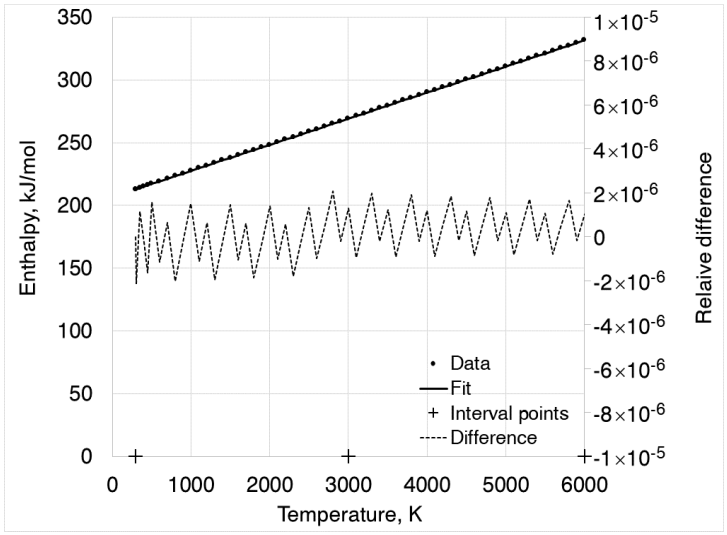


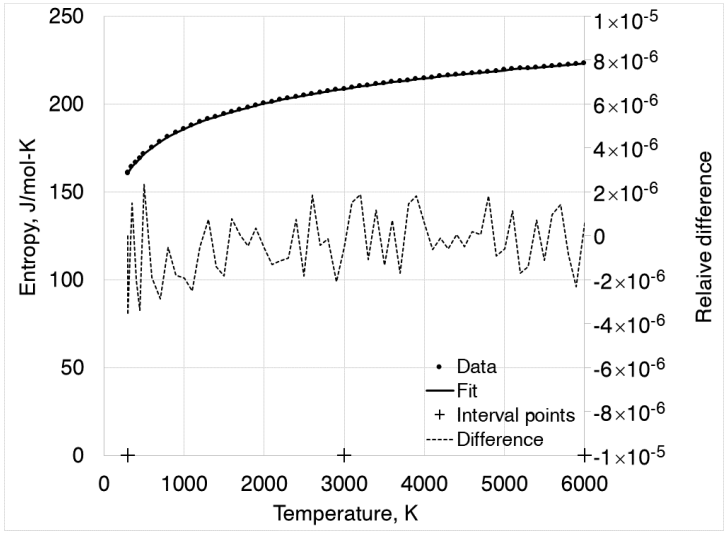


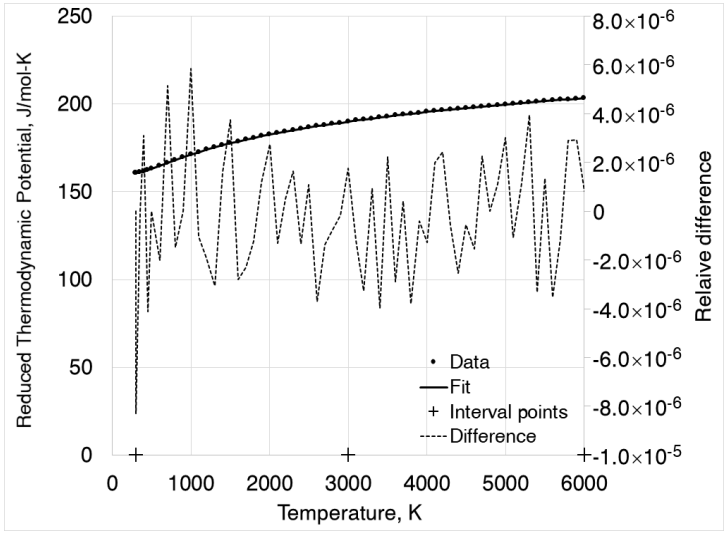
Copper, Ion (Cu^-) – Cu1-(g)

MW = 63.547 g/mol – $\Delta_f H^o$(298.15 K) = 212.92 kJ/mol – S^o(298.15 K) = 160.634 J/mol-K – p^o = 0.1 MPa



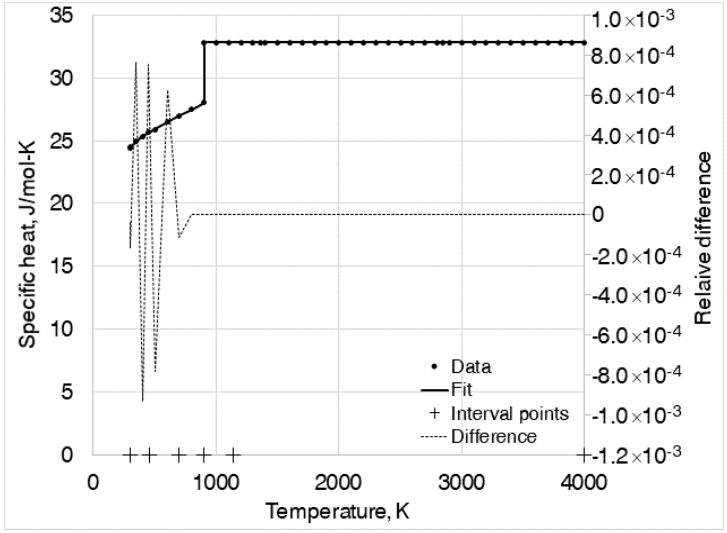


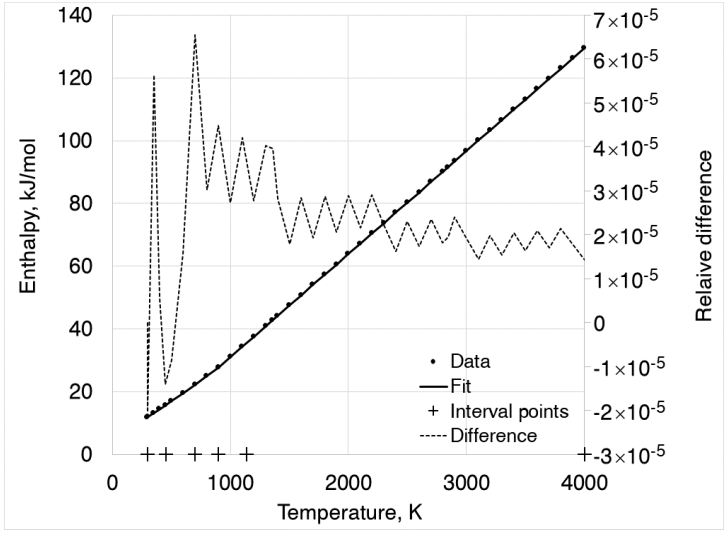


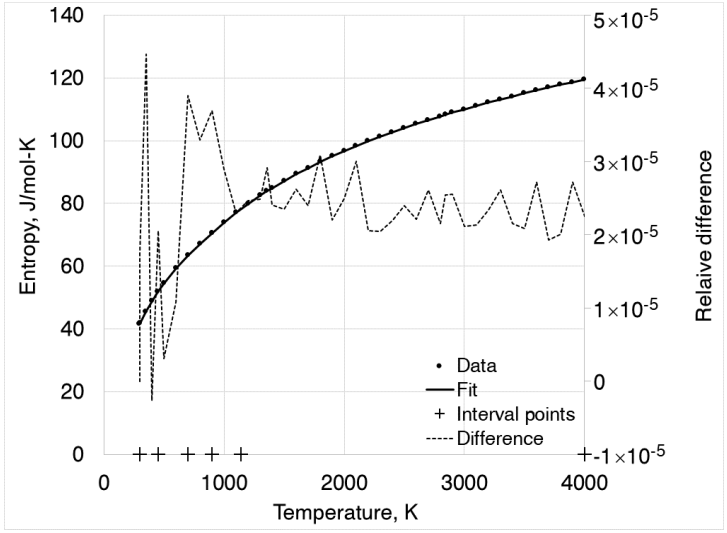


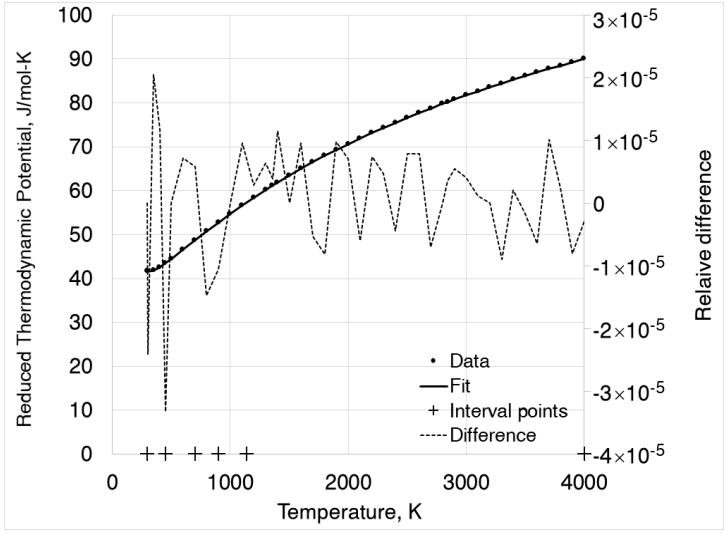
Copper (Cu) – Cu1(l)

MW = 63.546 g/mol – $\Delta_f H^o$(298.15 K) = 11.857 kJ/mol – S^o(298.15 K) = 41.62 J/mol-K – p^o = 0.1 MPa



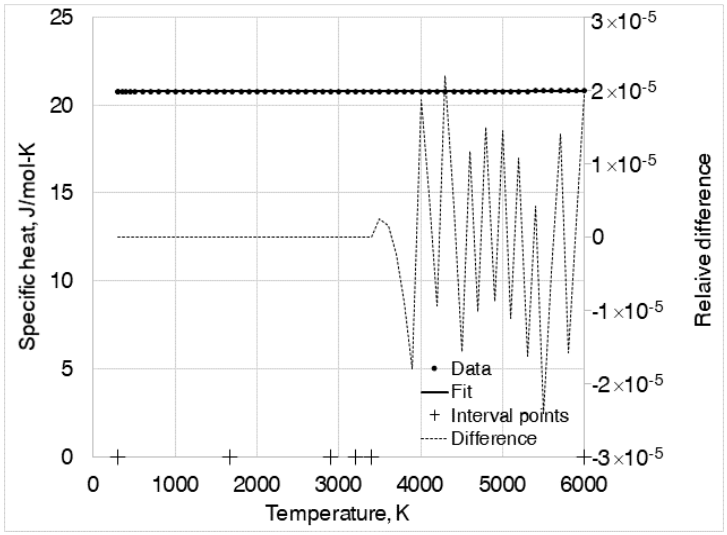


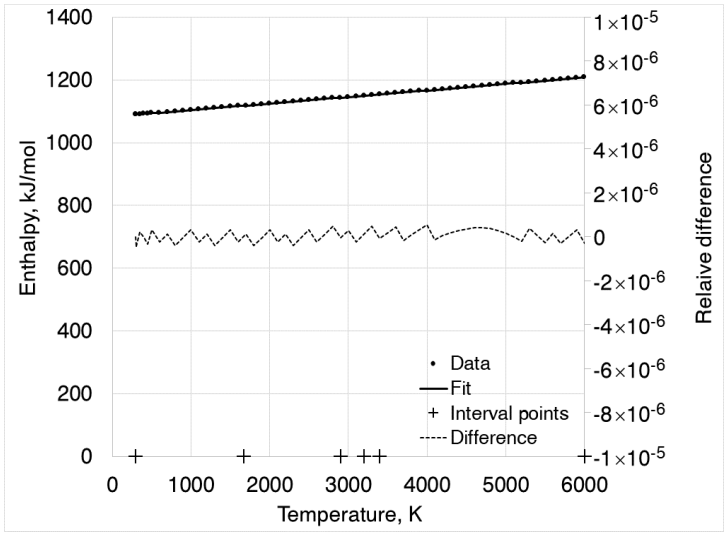


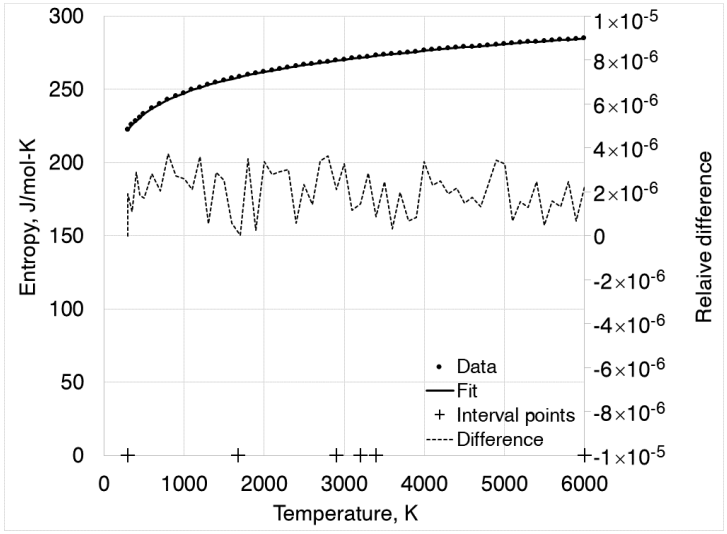


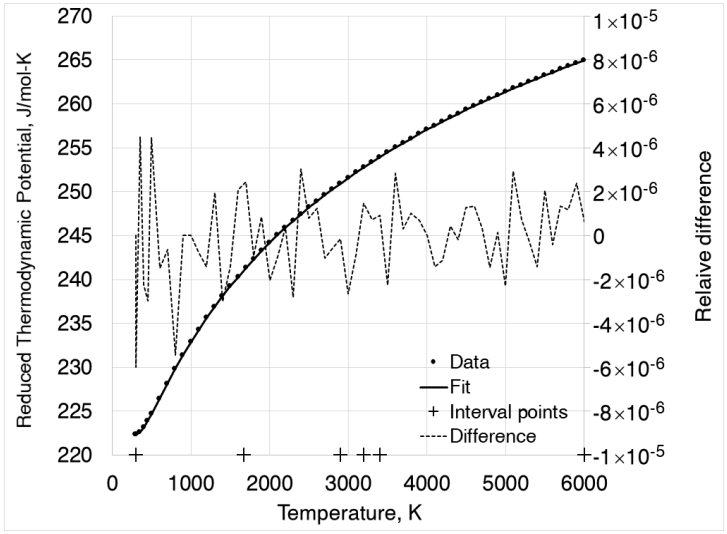
Copper, Ion (Cu^{+}) – Cu1+(g)

MW = 63.545 g/mol – $\Delta_f H^o$(298.15 K) = 1089.277 kJ/mol – S^o(298.15 K) = 222.311 J/mol-K – p^o = 0.1 MPa



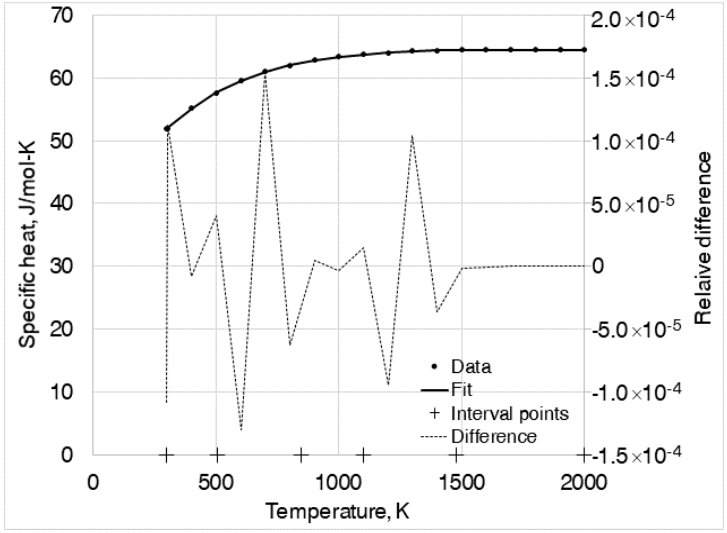


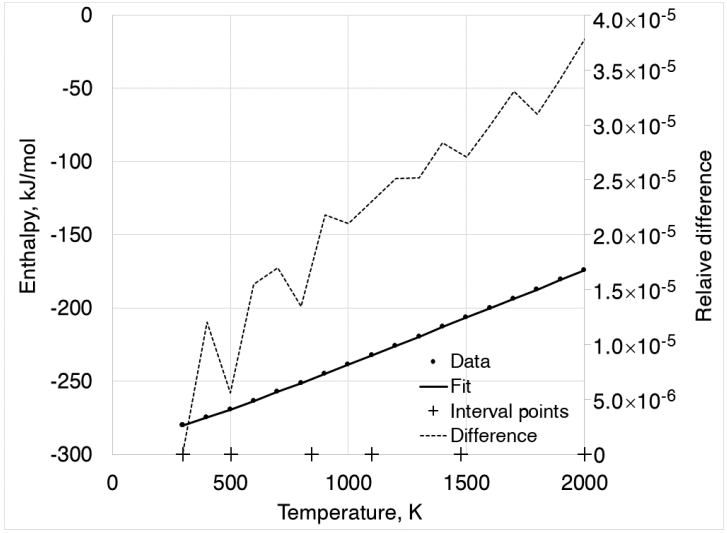


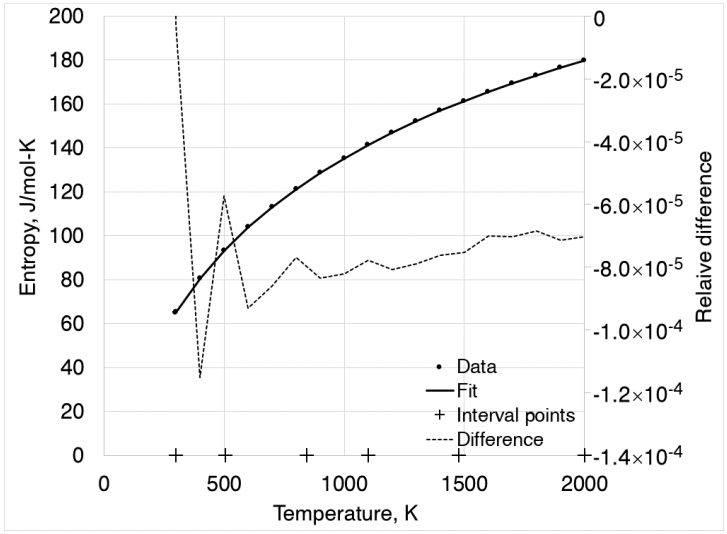


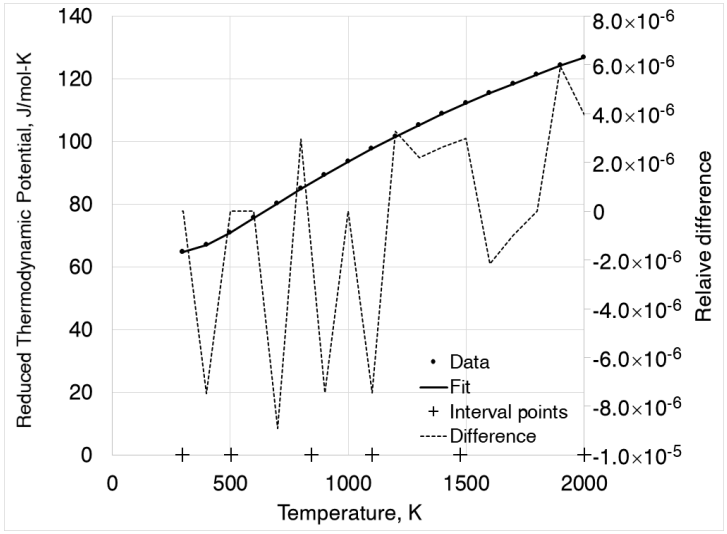
Copper Fluoride (CuF) – Cu1F1(cr)

MW = 82.544 g/mol – $\Delta_f H^o$(298.15 K) = -280.328 kJ/mol – S^o(298.15 K) = 64.852 J/mol-K – p^o = 0.1 MPa



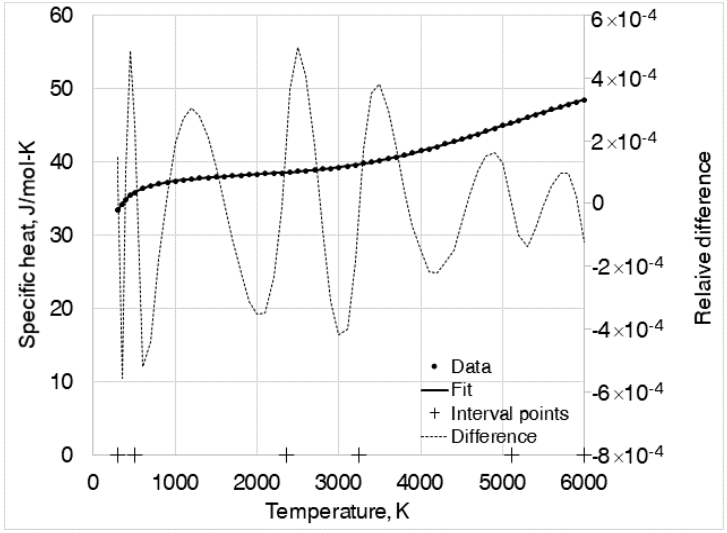


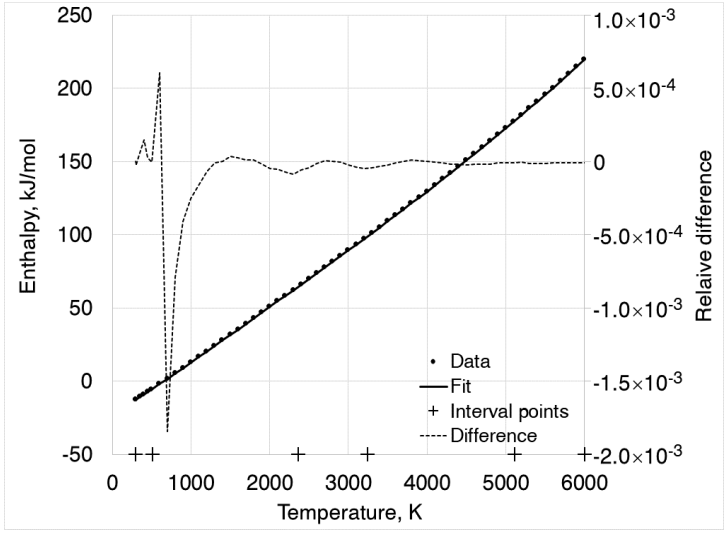


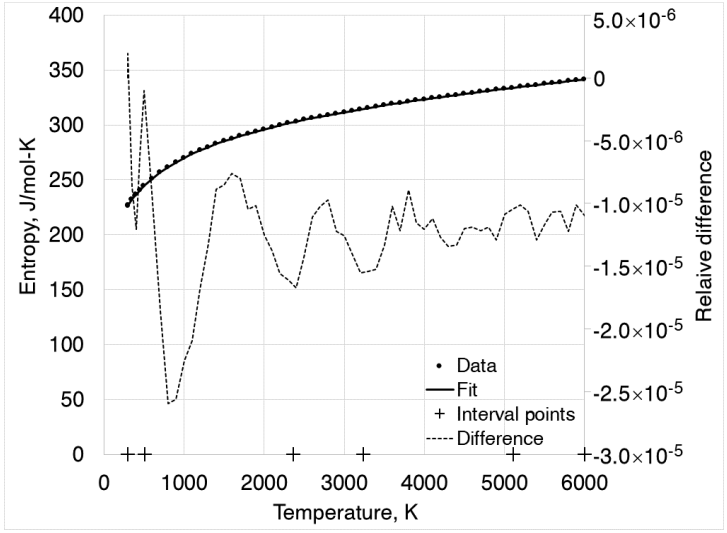


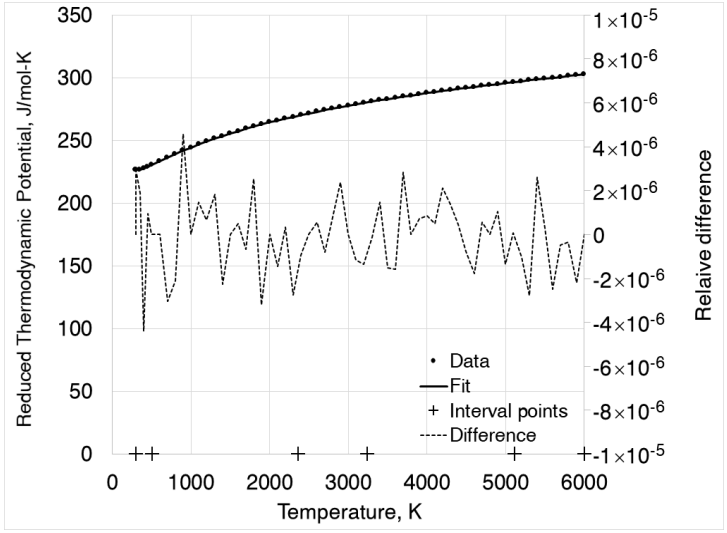
Copper Fluoride (CuF) – Cu1F1(g)

MW = 82.544 g/mol – $\Delta_f H^o$(298.15 K) = -12.552 kJ/mol – S^o(298.15 K) = 226.495 J/mol-K – p^o = 0.1 MPa



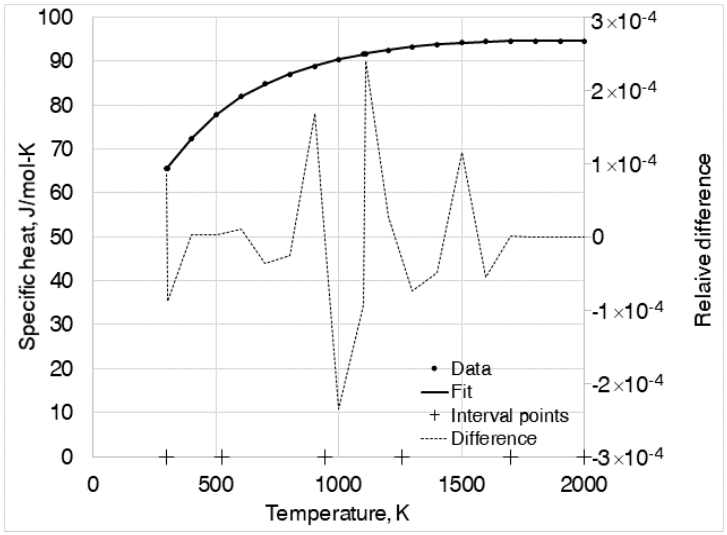


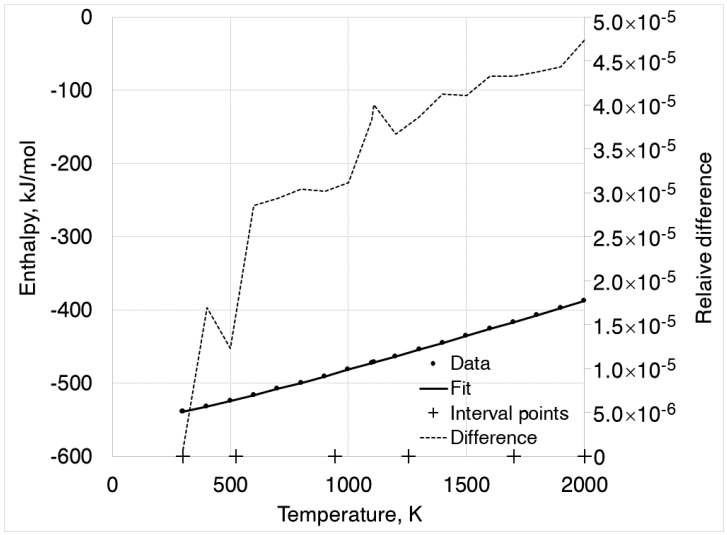


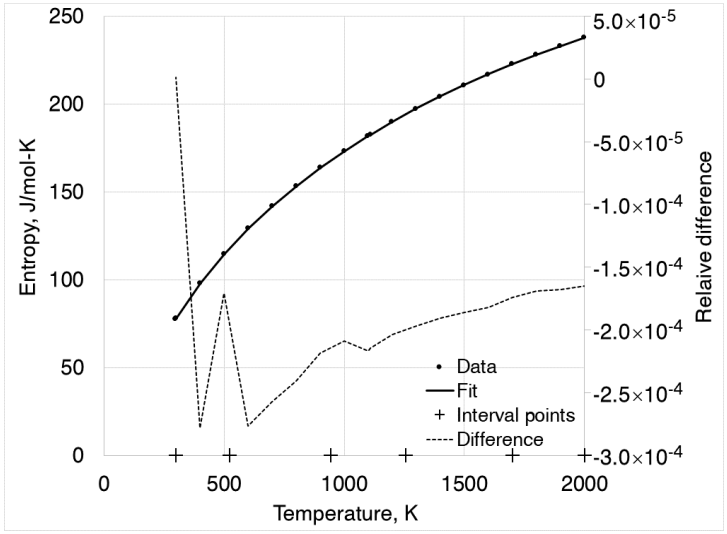


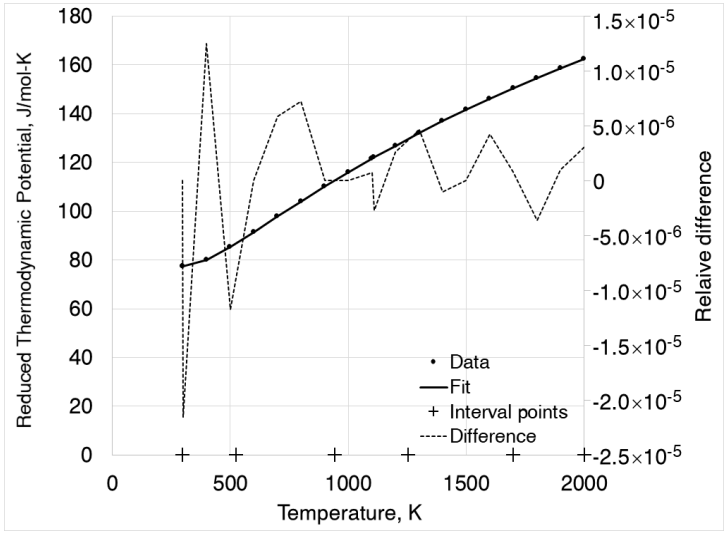
Copper Fluoride (CuF_2) – Cu1F2(cr)

MW = 101.543 g/mol – $\Delta_f H^o$(298.15 K) = -538.899 kJ/mol – S^o(298.15 K) = 77.452 J/mol-K – p^o = 0.1 MPa



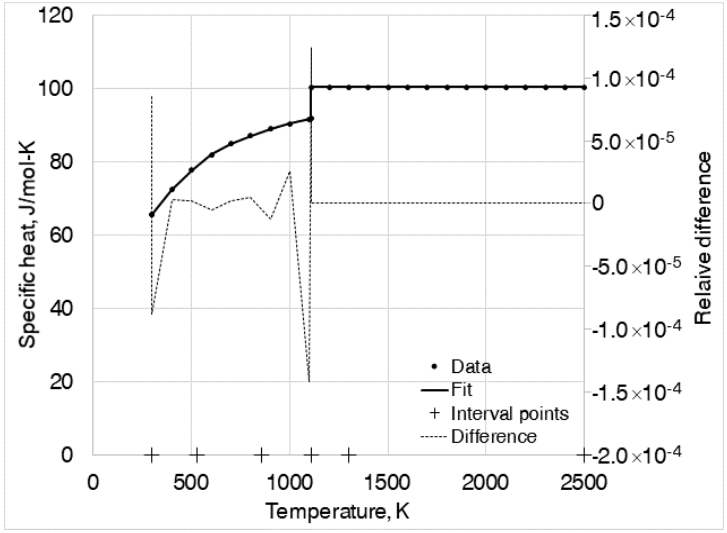


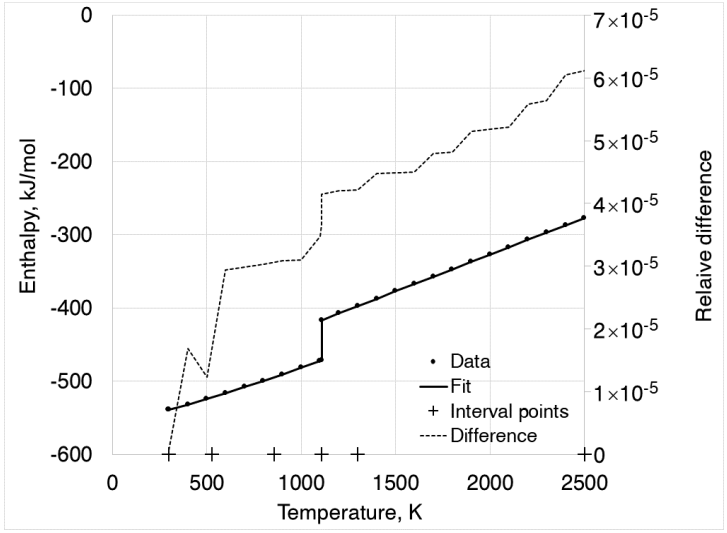


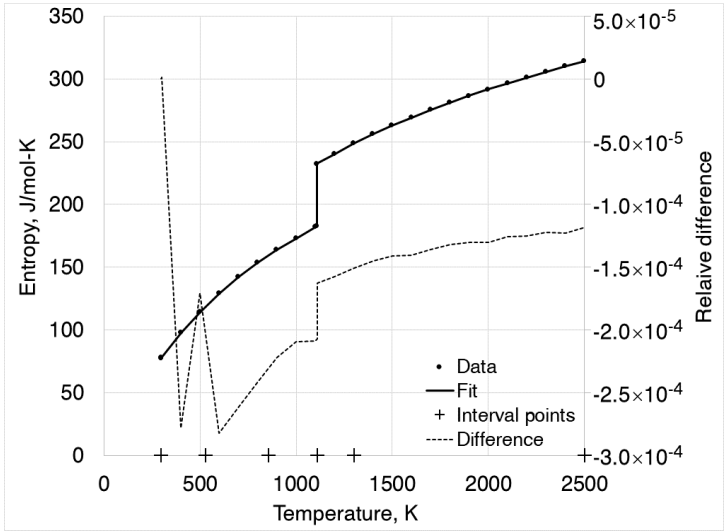


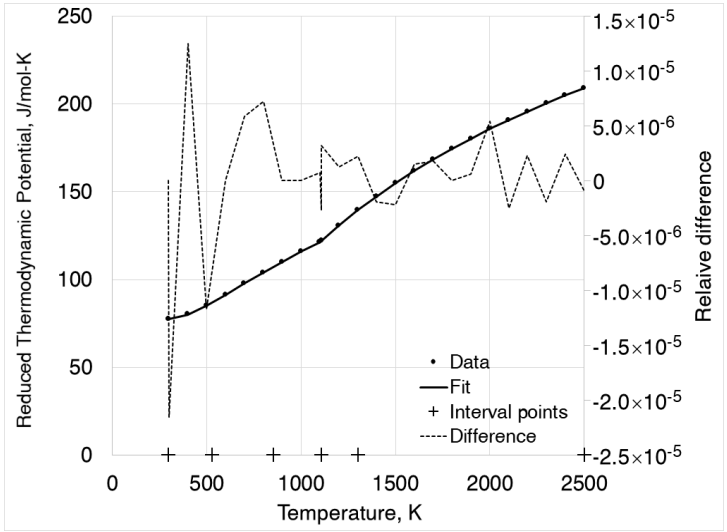
Copper Fluoride (CuF_2) – Cu1F2(cr,l)

MW = 101.543 g/mol – $\Delta_f H^o$(298.15 K) = -538.899 kJ/mol – S^o(298.15 K) = 77.452 J/mol-K – p^o = 0.1 MPa



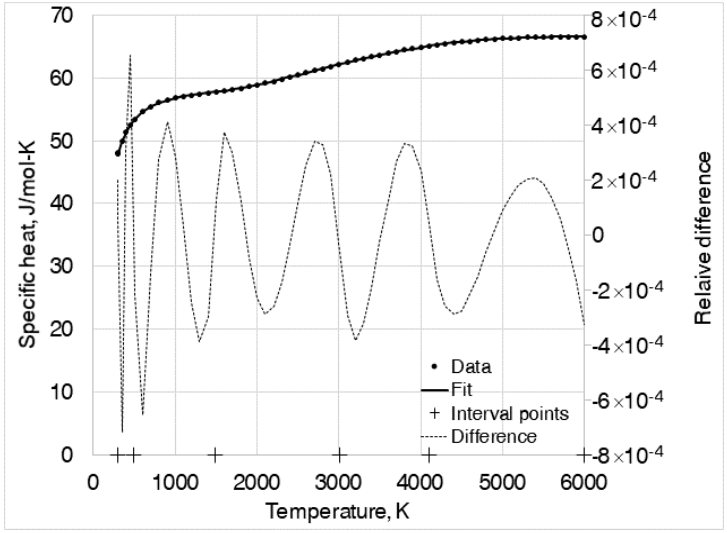


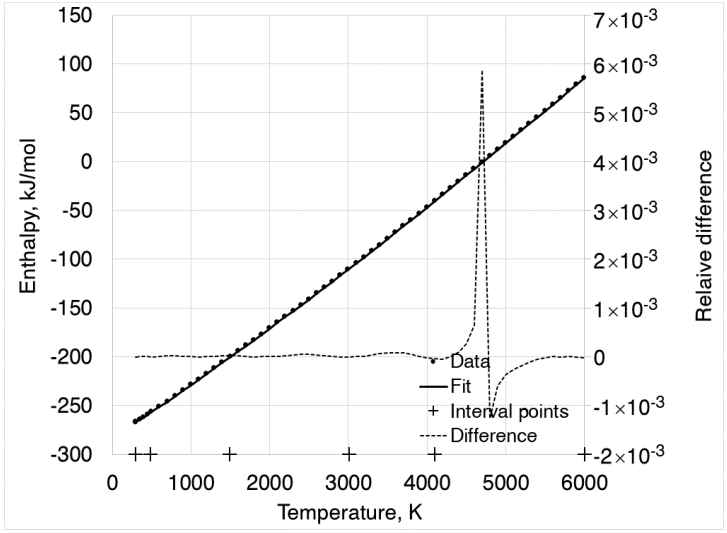


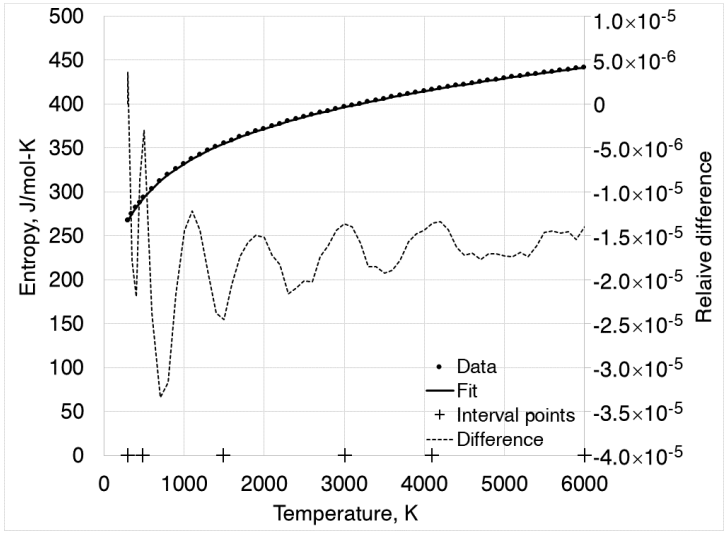


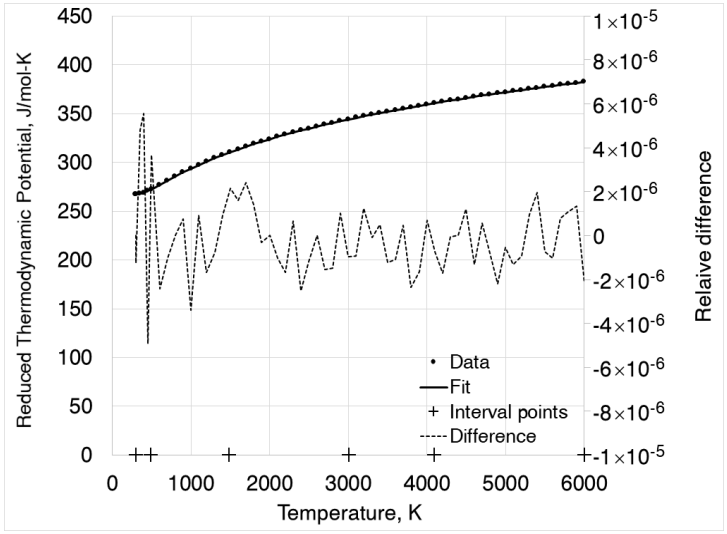
Copper Fluoride (CuF_2) – Cu1F2(g)

MW = 101.543 g/mol – $\Delta_f H^o$(298.15 K) = -266.939 kJ/mol – S^o(298.15 K) = 267.086 J/mol-K – p^o = 0.1 MPa



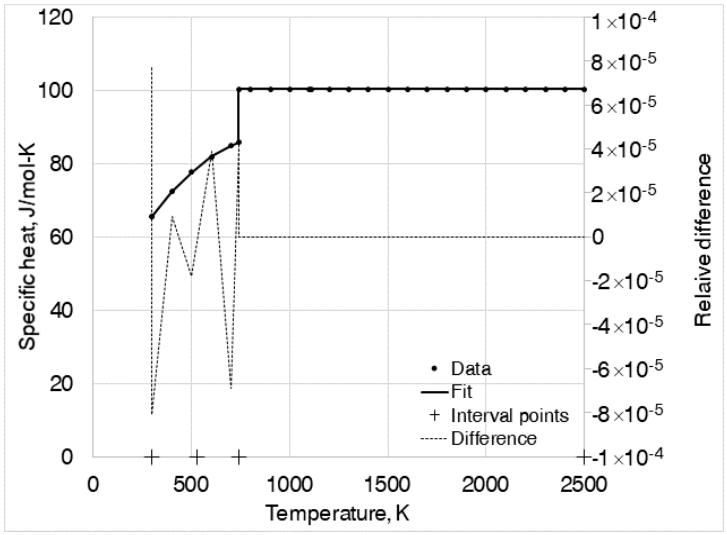


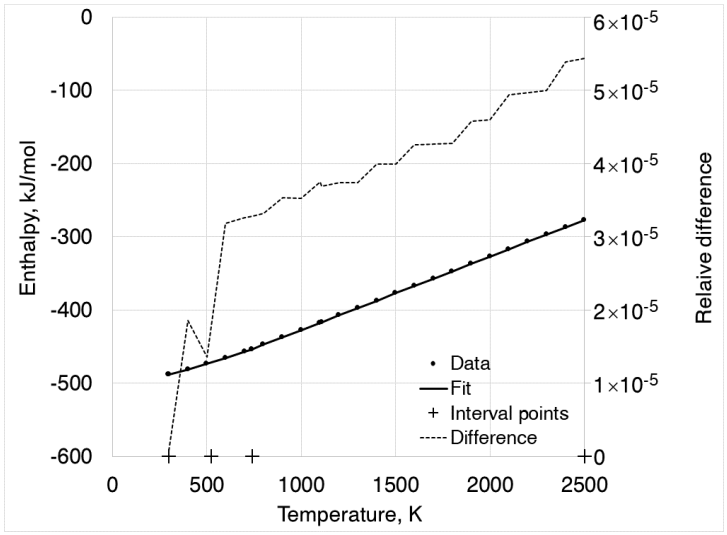


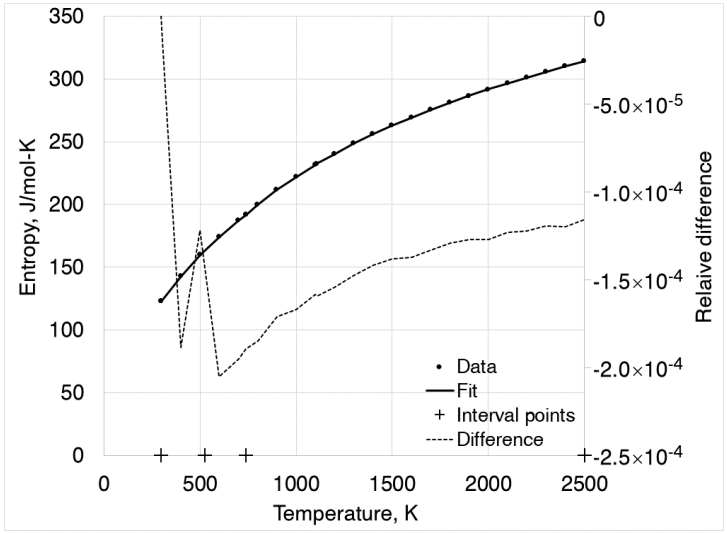


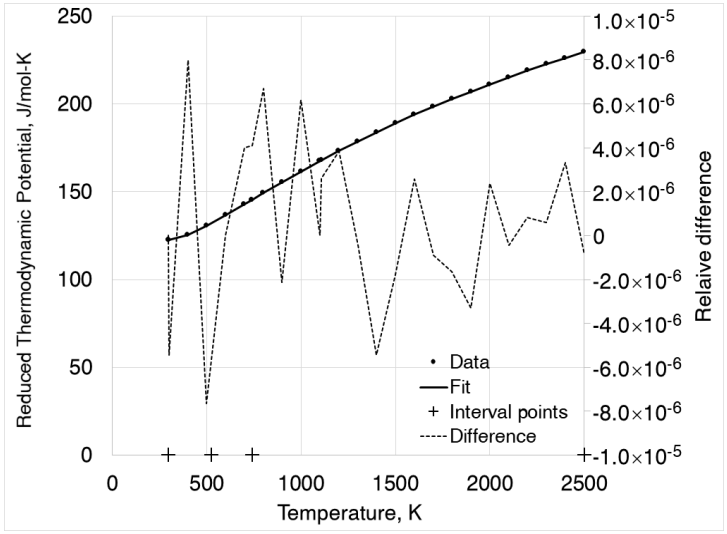
Copper Fluoride (CuF_2) – Cu1F2(l)

MW = 101.543 g/mol – $\Delta_f H^o$(298.15 K) = -487.854 kJ/mol – S^o(298.15 K) = 122.585 J/mol-K – p^o = 0.1 MPa



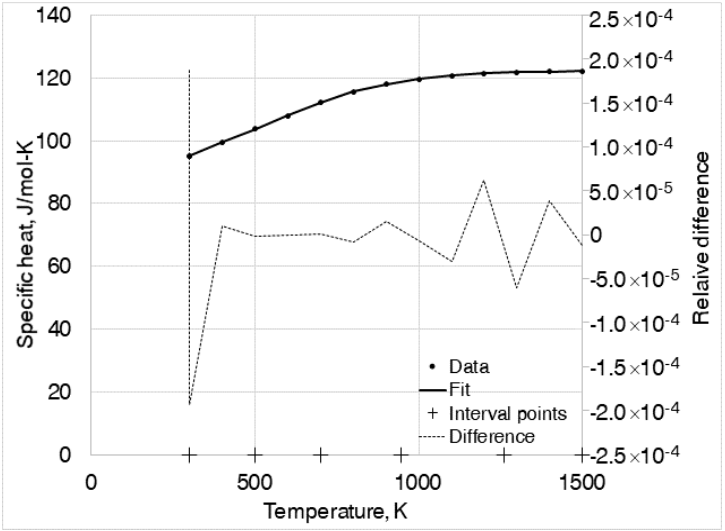


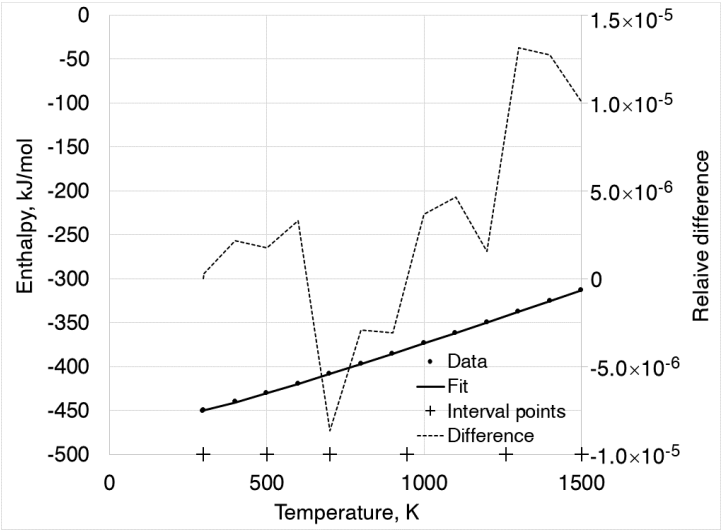


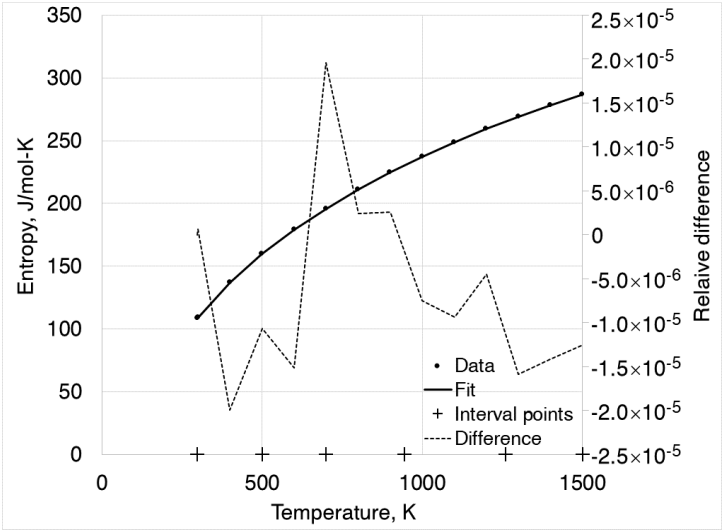


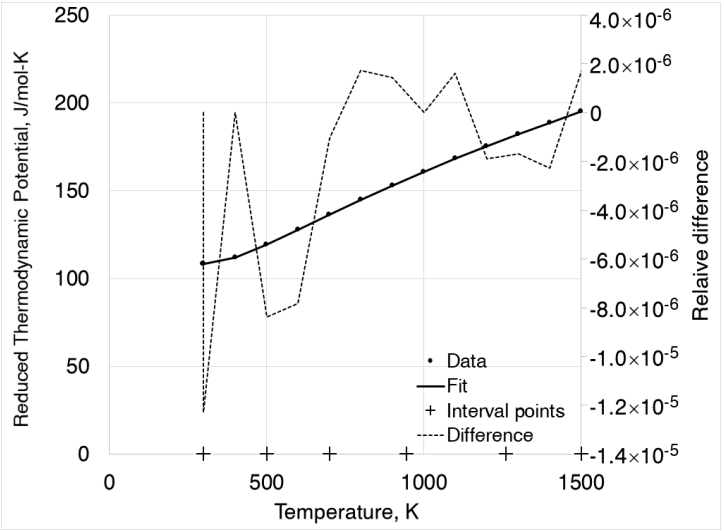
Copper Hydroxide ($Cu(OH)_2$) – Cu1H2O2(cr)

MW = 97.561 g/mol – $\Delta_f H^o$(298.15 K) = -450.366 kJ/mol – S^o(298.15 K) = 108.366 J/mol-K – p^o = 0.1 MPa



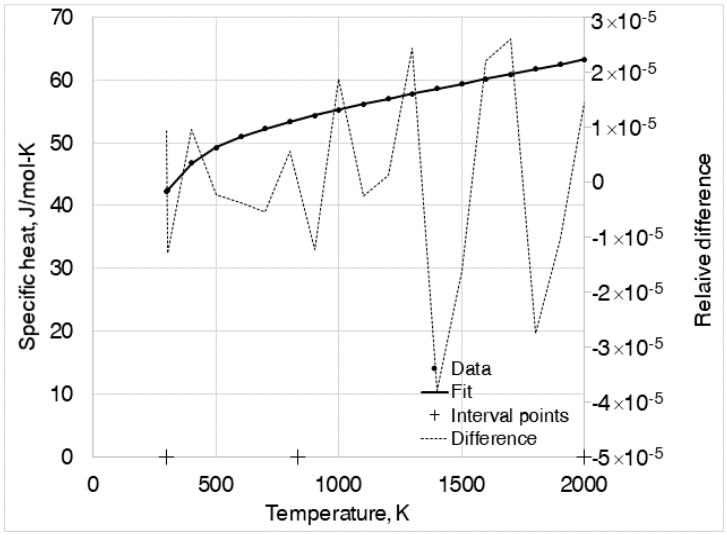


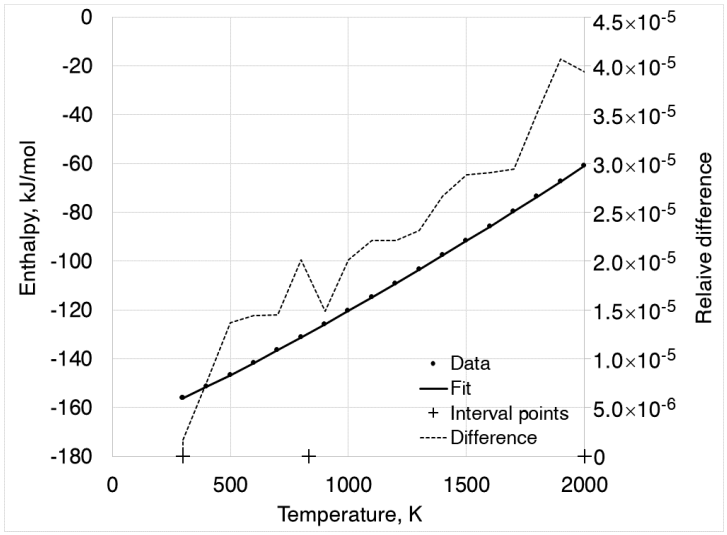


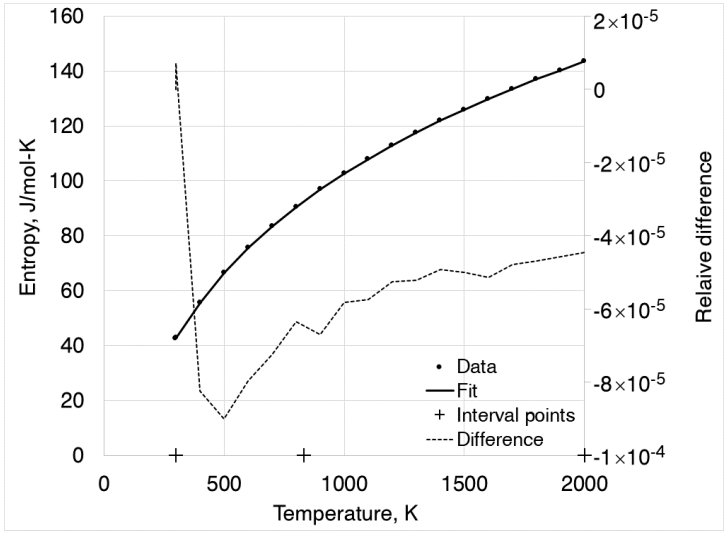


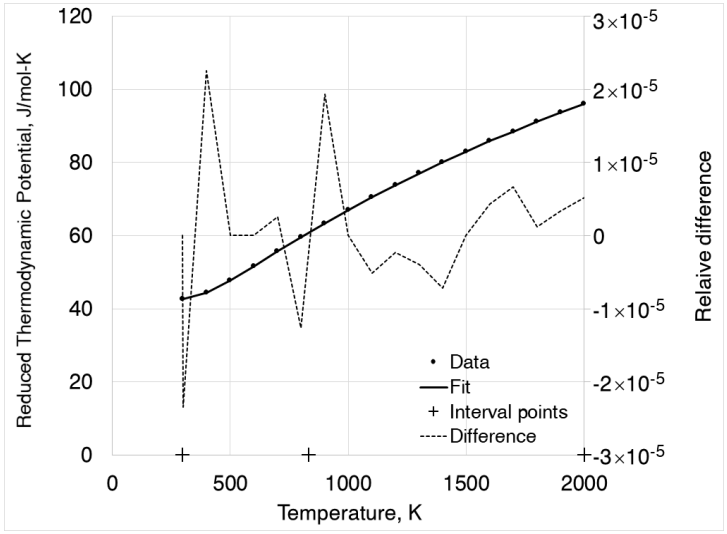
Copper Oxide (CuO) – Cu1O1(cr)

MW = 79.545 g/mol – $\Delta_f H^o$(298.15 K) = -156.063 kJ/mol – S^o(298.15 K) = 42.594 J/mol-K – p^o = 0.1 MPa



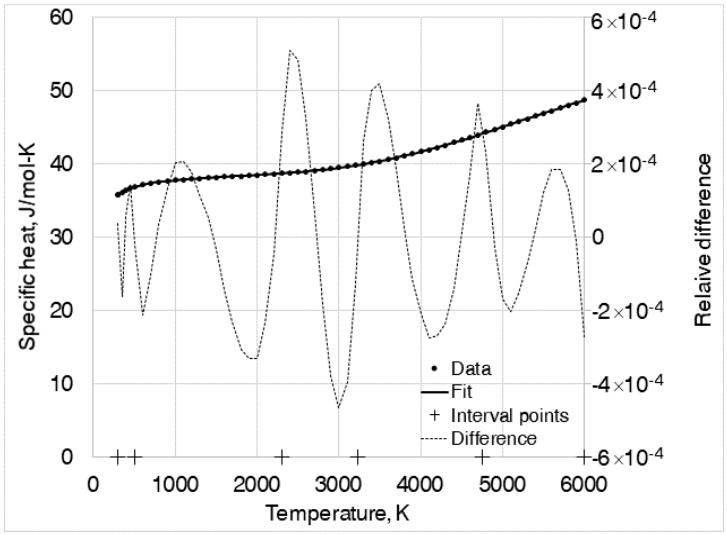


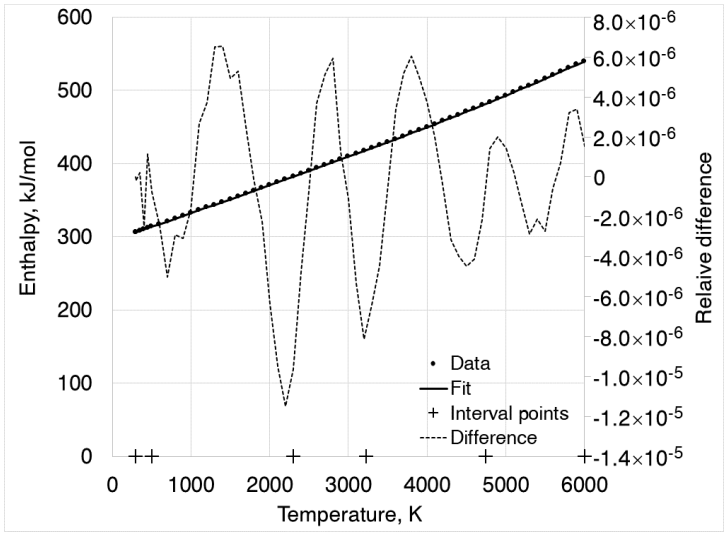


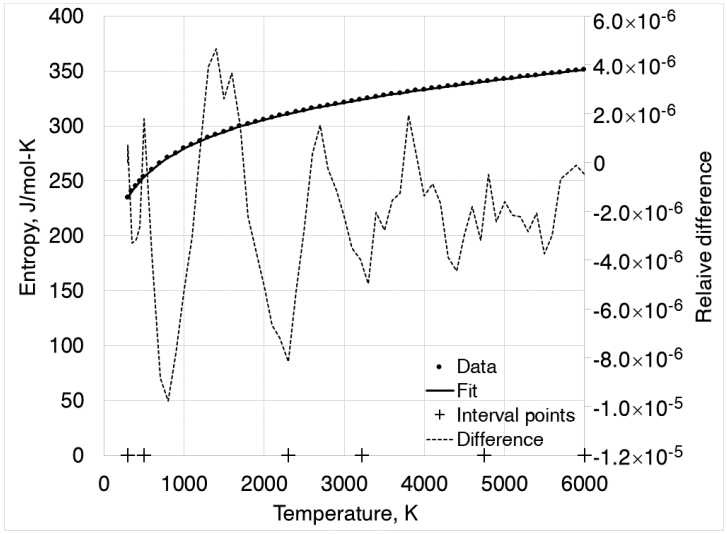


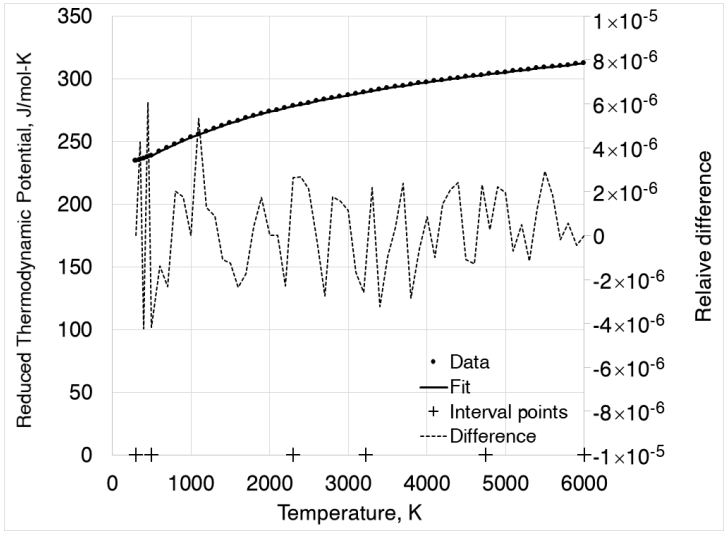
Copper Oxide (CuO) – Cu1O1(g)

MW = 79.545 g/mol – $\Delta_f H^o$(298.15 K) = 306.269 kJ/mol – S^o(298.15 K) = 234.617 J/mol-K – p^o = 0.1 MPa



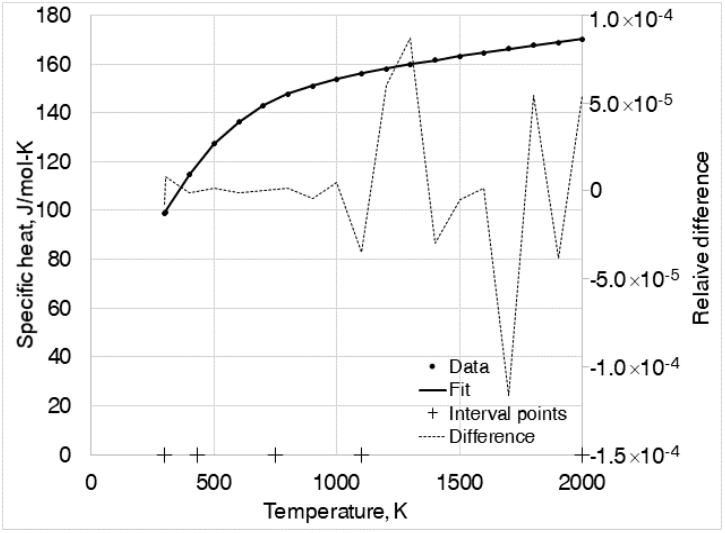


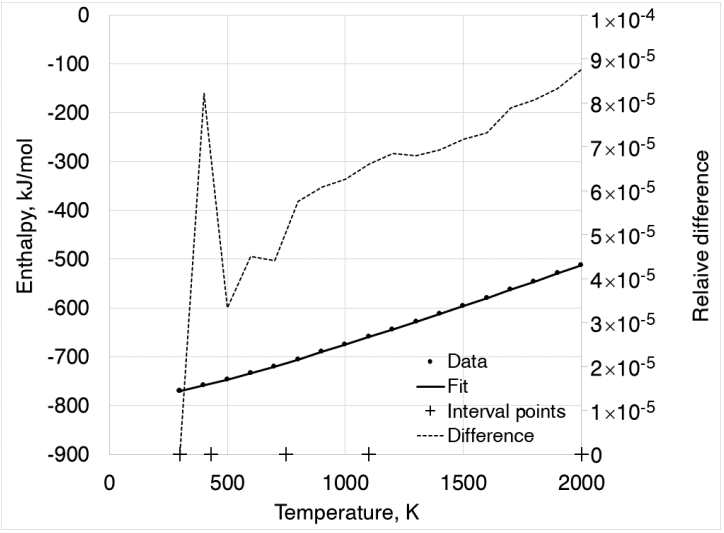


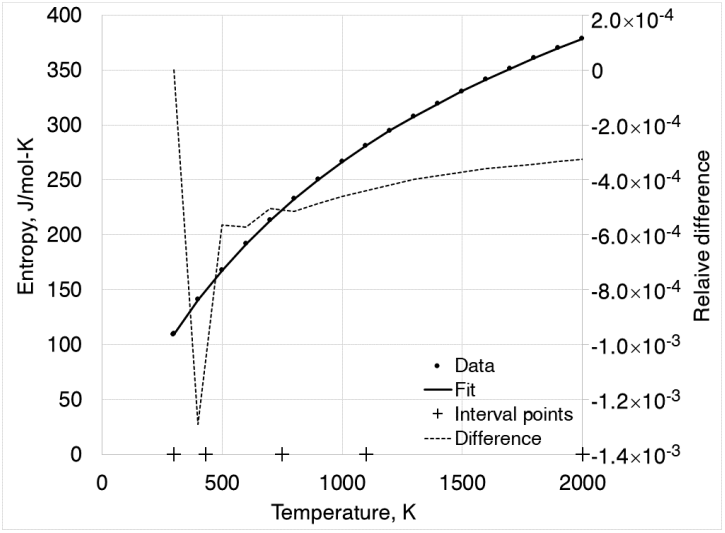


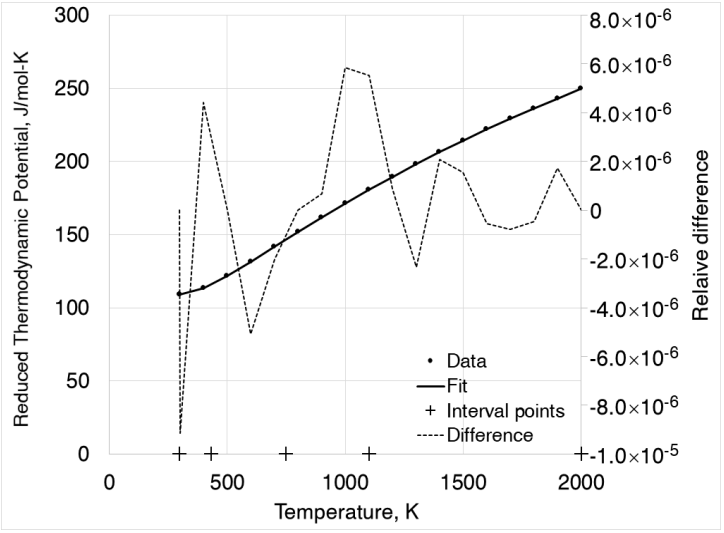
Copper Sulfate ($CuSO_4$) – Cu1O4S1(cr)

MW = 159.6 g/mol – $\Delta_f H^o$(298.15 K) = -769.982 kJ/mol – S^o(298.15 K) = 109.254 J/mol-K – p^o = 0.1 MPa



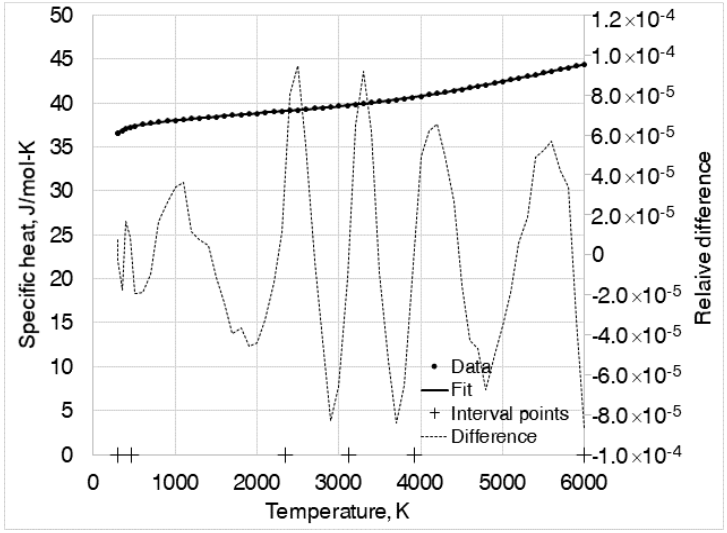


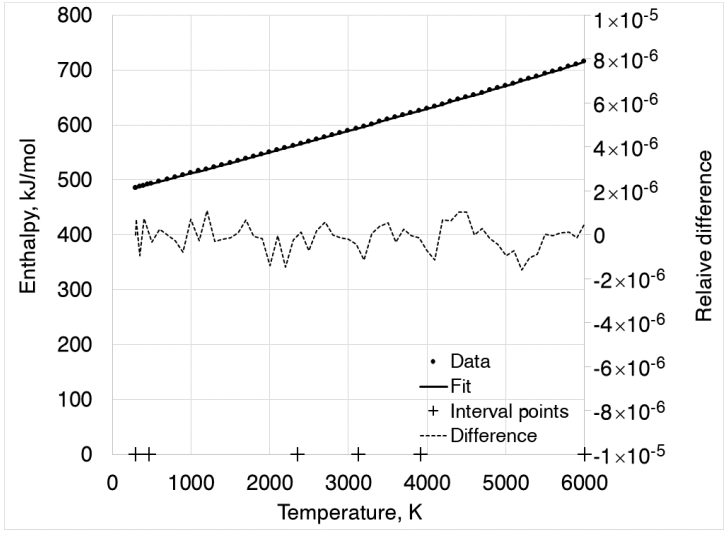


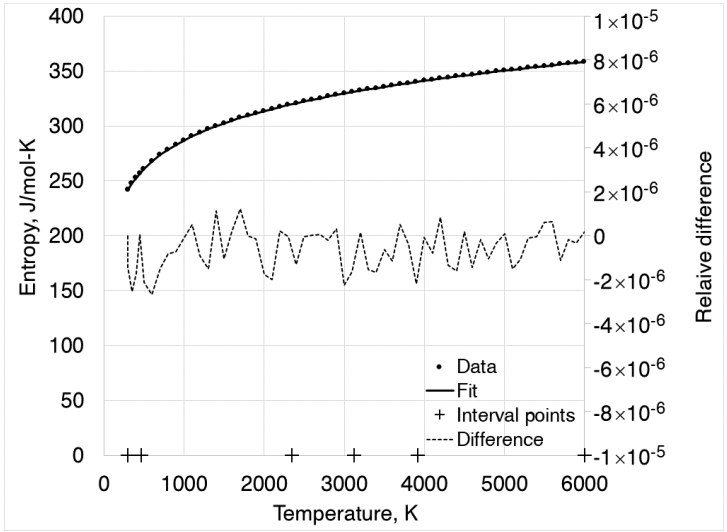


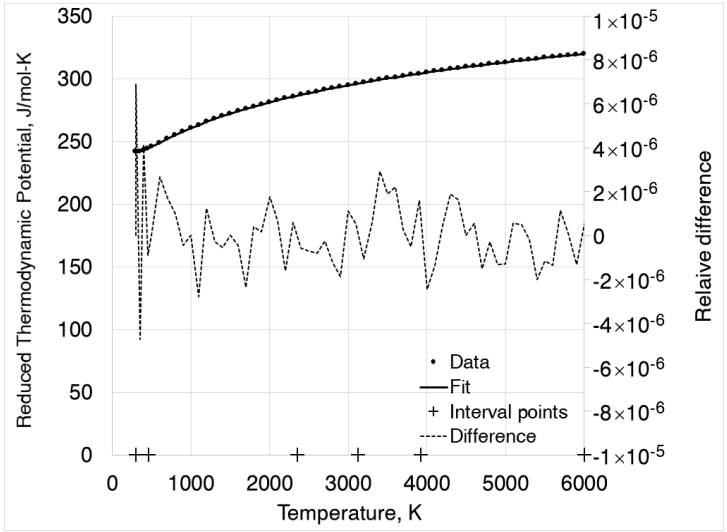
Copper (Cu_2) – Cu2(g)

MW = 127.092 g/mol – $\Delta_f H^o$(298.15 K) = 485.344 kJ/mol – S^o(298.15 K) = 241.721 J/mol-K – p^o = 0.1 MPa



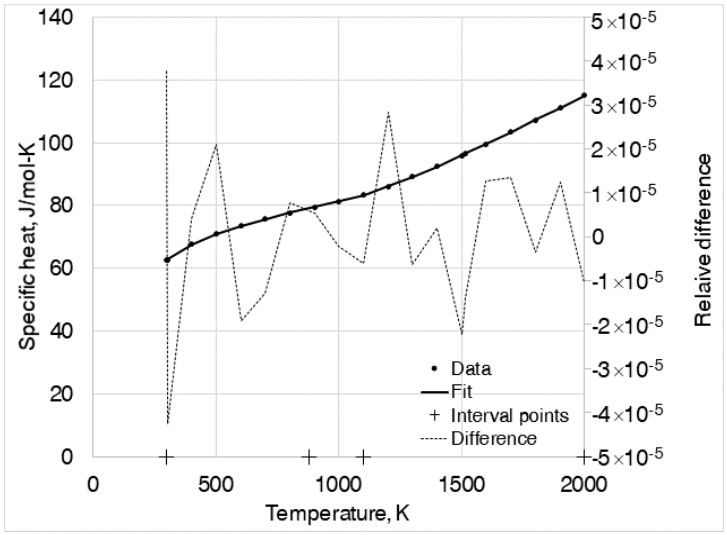


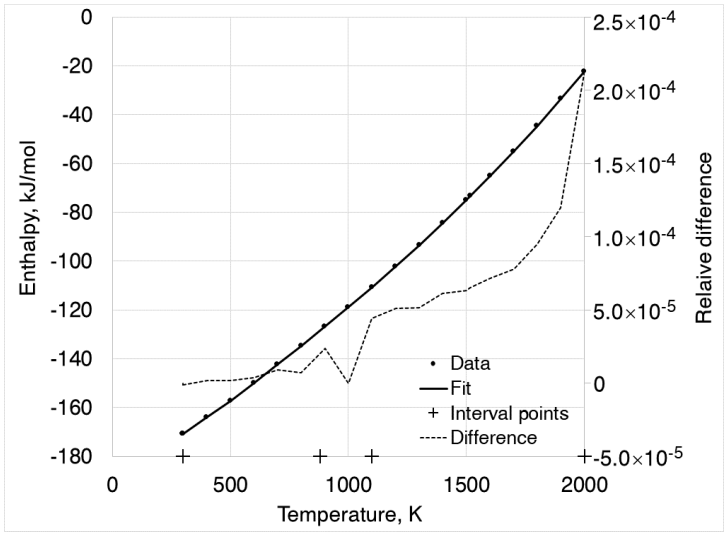


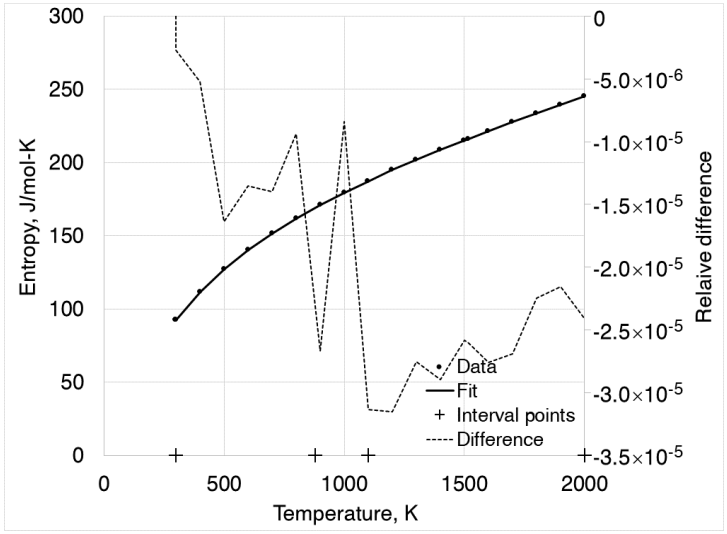


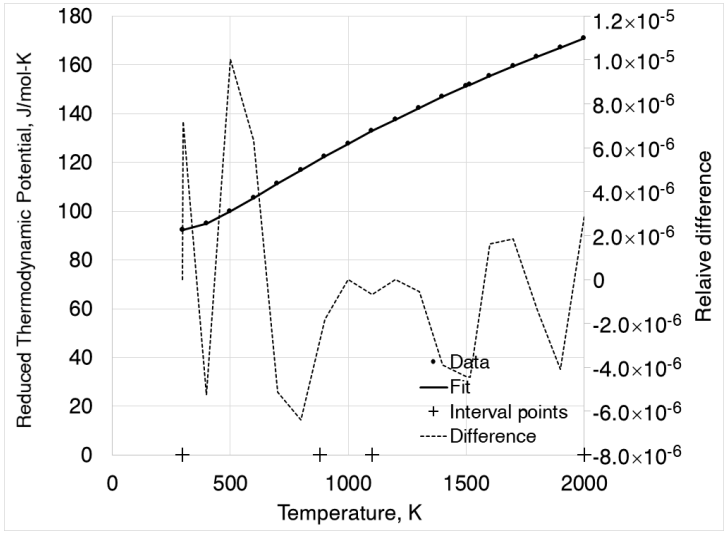
Copper Oxide (Cu_2O) – Cu2O1(cr)

MW = 143.091 g/mol – $\Delta_f H^o$(298.15 K) = -170.707 kJ/mol – S^o(298.15 K) = 92.36 J/mol-K – p^o = 0.1 MPa



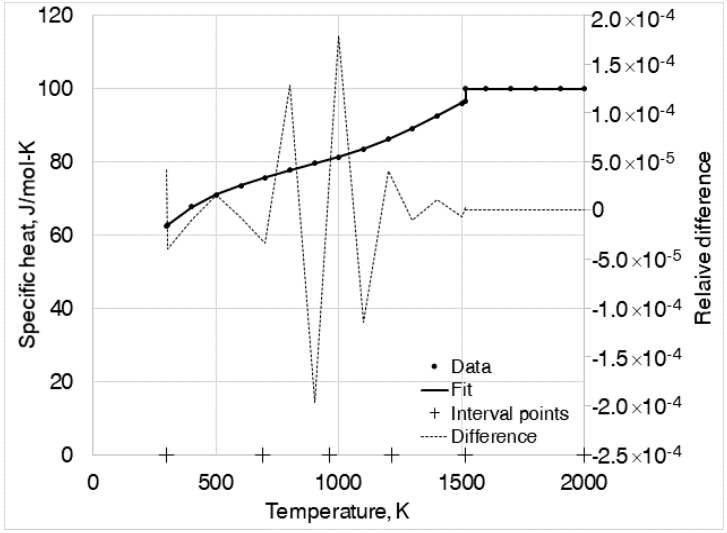


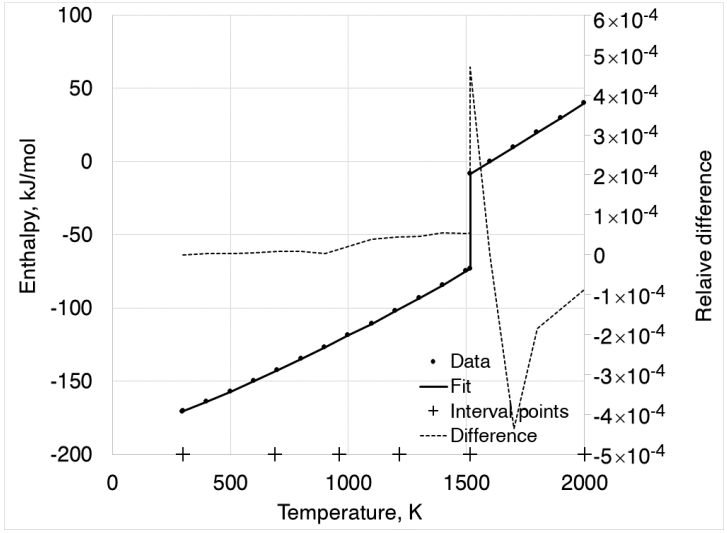


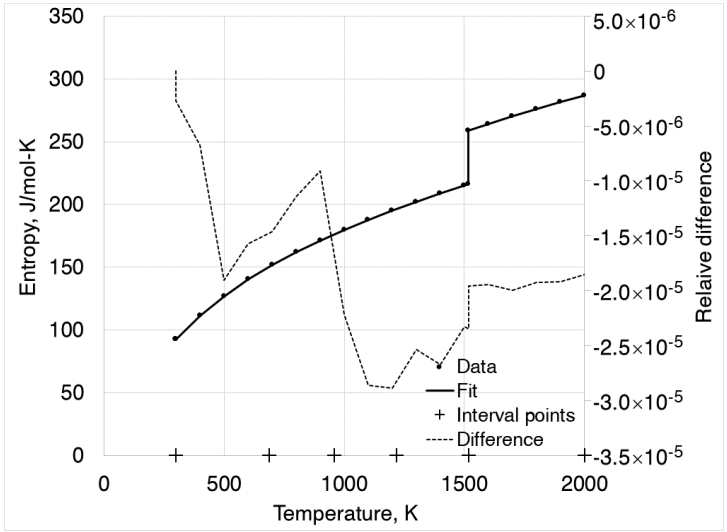


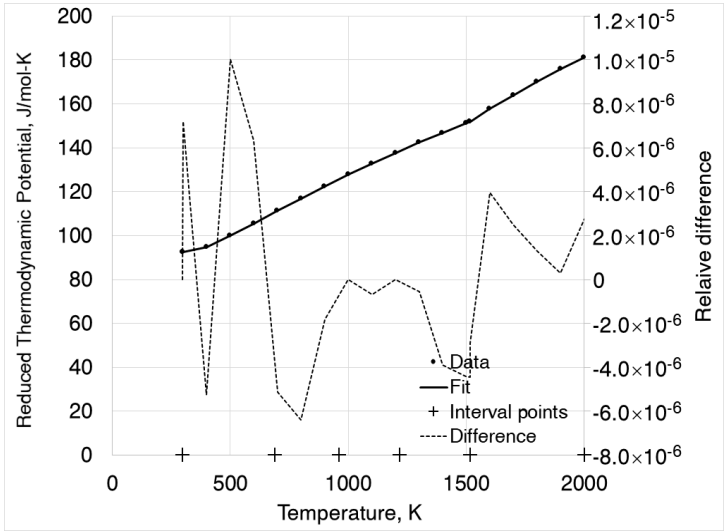
Copper Oxide (Cu_2O) – Cu2O1(cr,l)

MW = 143.091 g/mol – $\Delta_f H^o$(298.15 K) = -170.707 kJ/mol – S^o(298.15 K) = 92.36 J/mol-K – p^o = 0.1 MPa



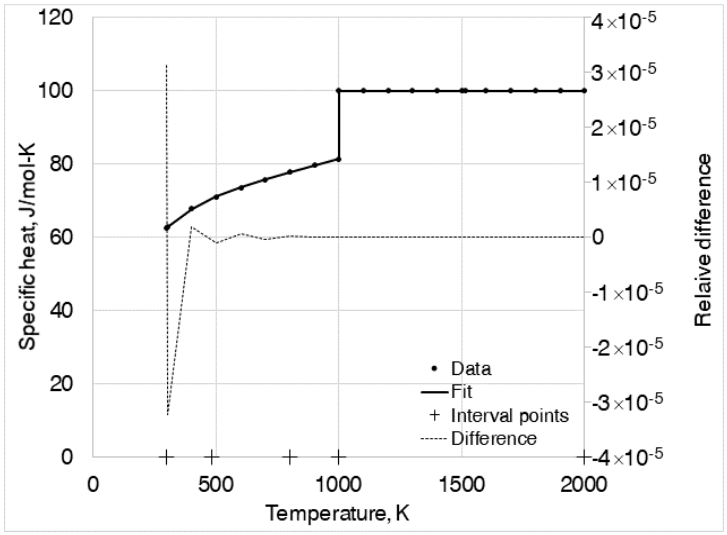


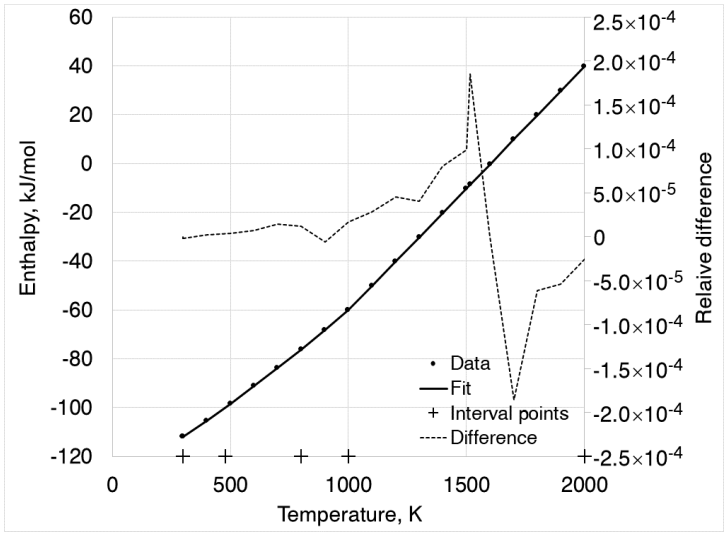


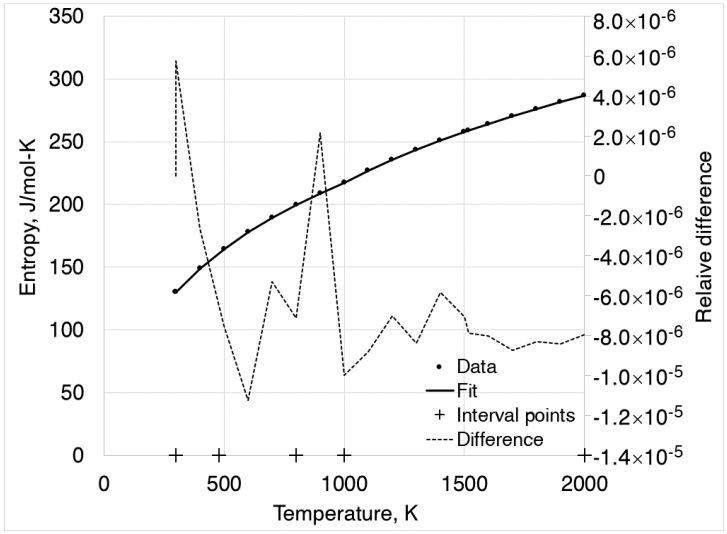


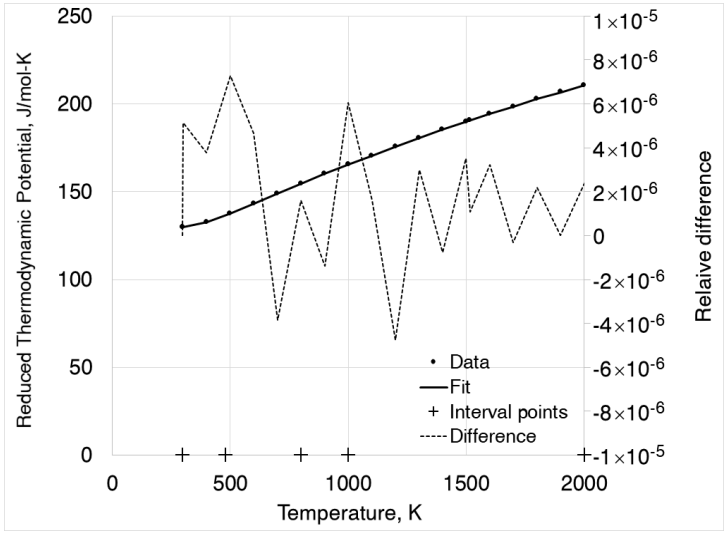
Copper Oxide (Cu_2O) – Cu2O1(l)

MW = 143.091 g/mol – $\Delta_f H^o$(298.15 K) = -111.995 kJ/mol – S^o(298.15 K) = 129.957 J/mol-K – p^o = 0.1 MPa



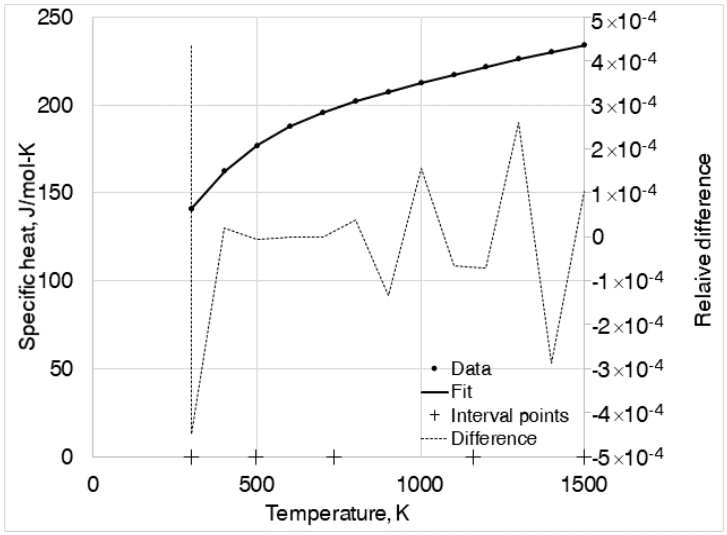


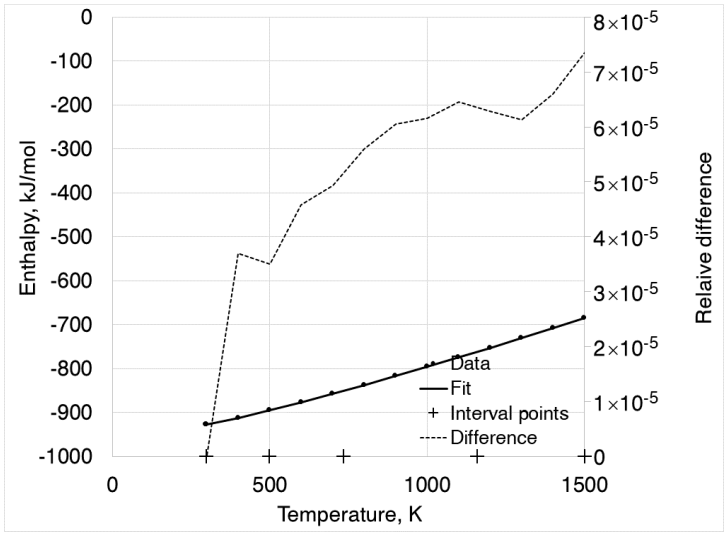


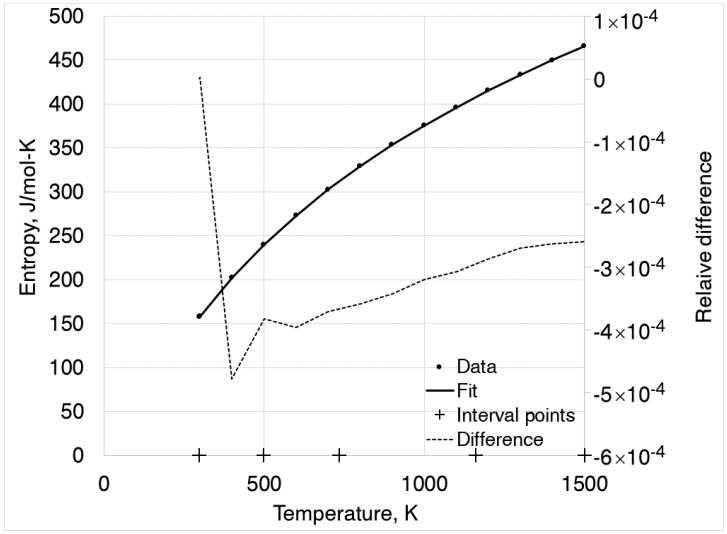


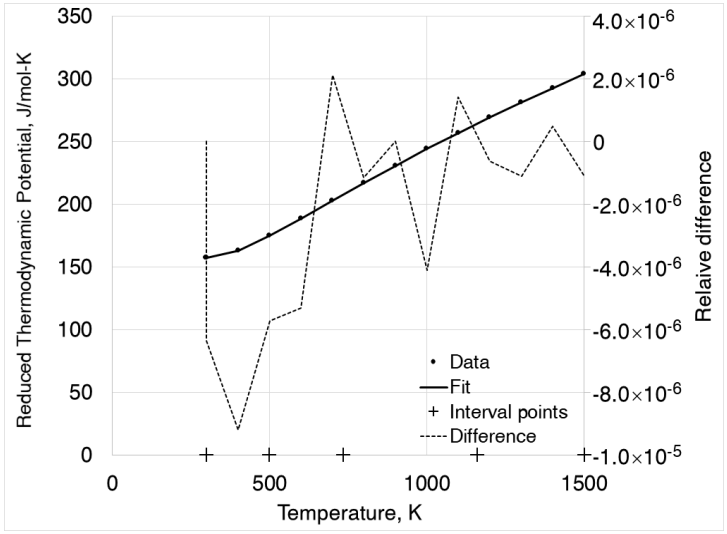
Copper Oxide Sulfate ($CuO.CuSO_4$) – Cu2O5S1(cr)

MW = 239.15 g/mol – $\Delta_f H^o$(298.15 K) = -927.593 kJ/mol – S^o(298.15 K) = 157.318 J/mol-K – p^o = 0.1 MPa



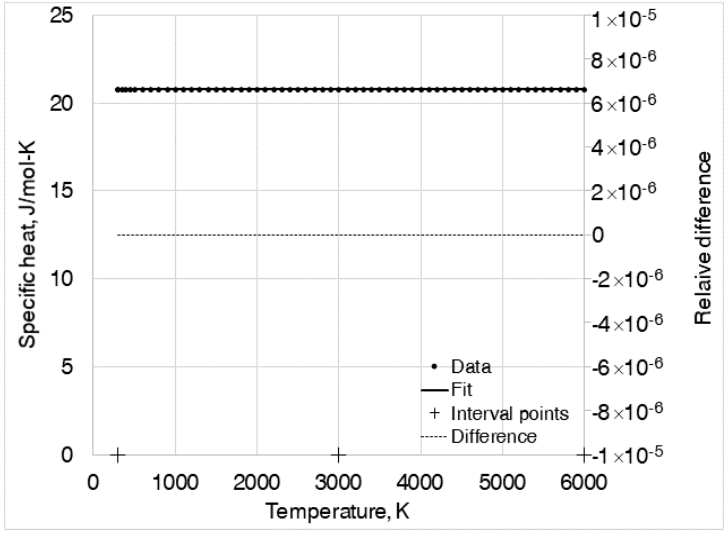


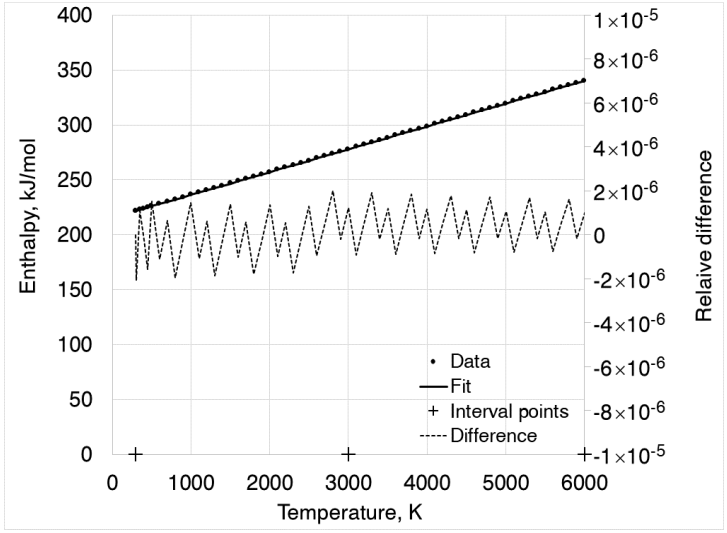


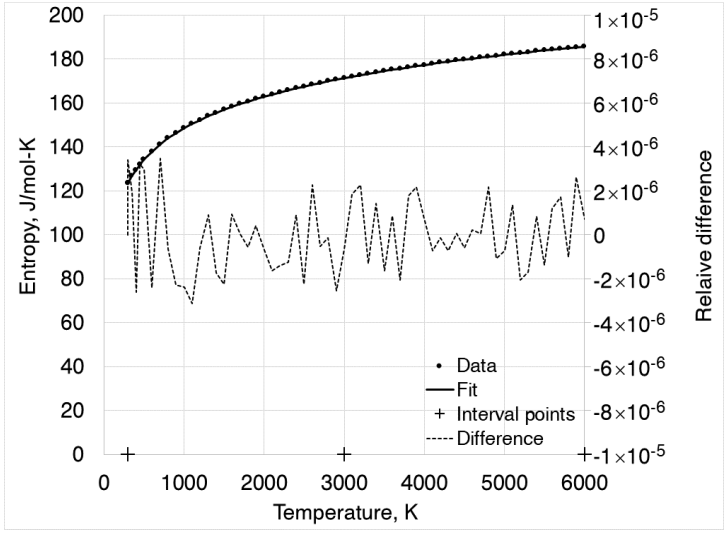


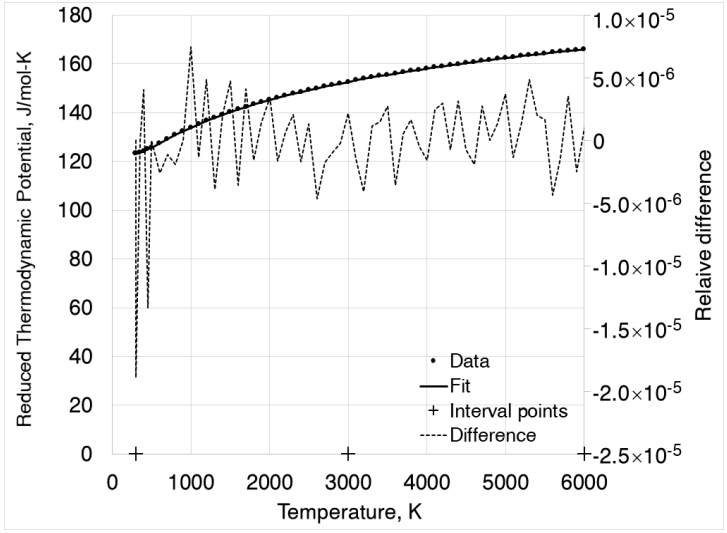
Deuterium (D) – D1(g)

MW = 294 g/mol – $\Delta_f H^o$(298.15 K) = 221.72 kJ/mol – S^o(298.15 K) = 123.35 J/mol-K – p^o = 0.1 MPa



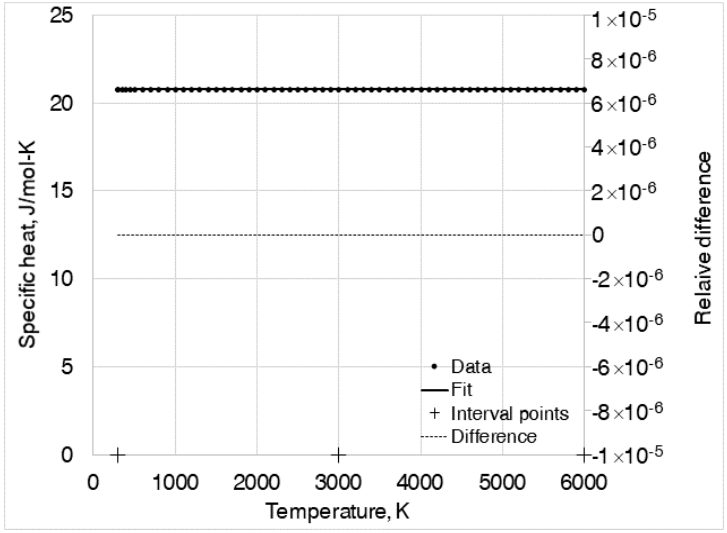


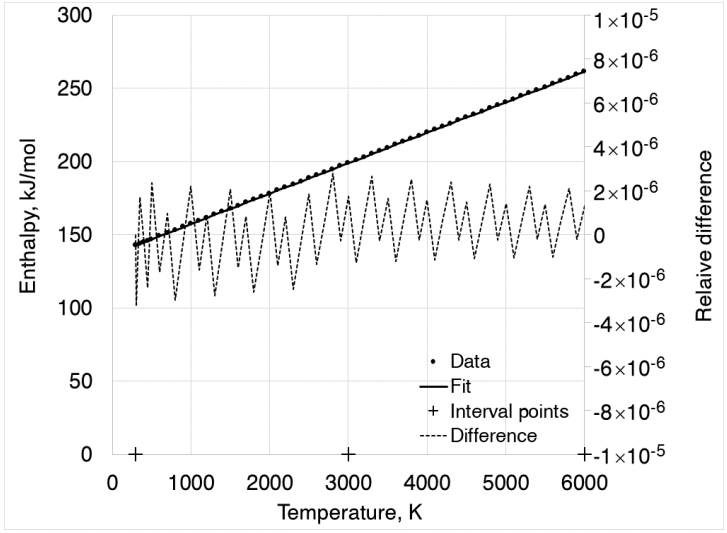


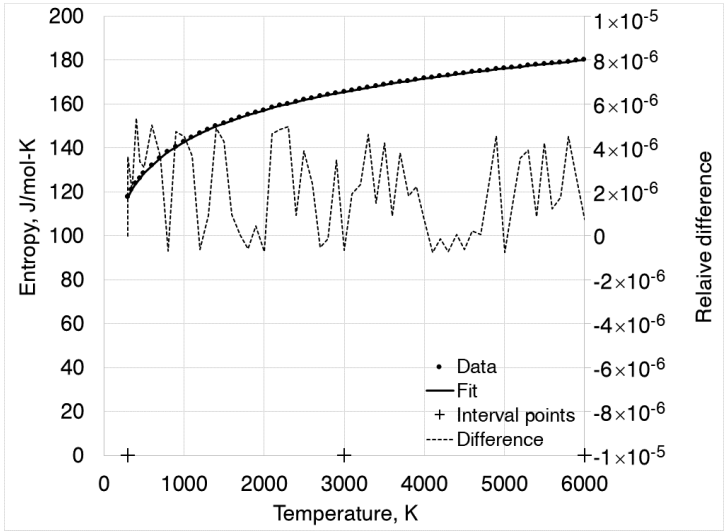


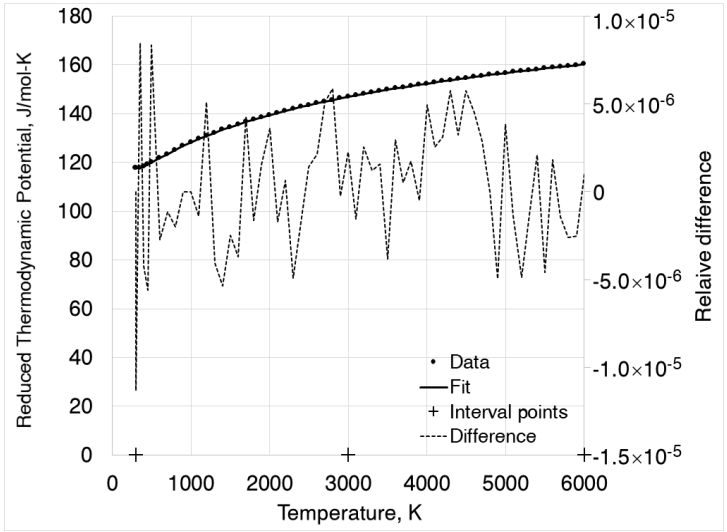
Deuterium, Ion (D^-) – D1-(g)

MW = 294 g/mol – $\Delta_f H^o$(298.15 K) = 142.753 kJ/mol – S^o(298.15 K) = 117.59 J/mol-K – p^o = 0.1 MPa



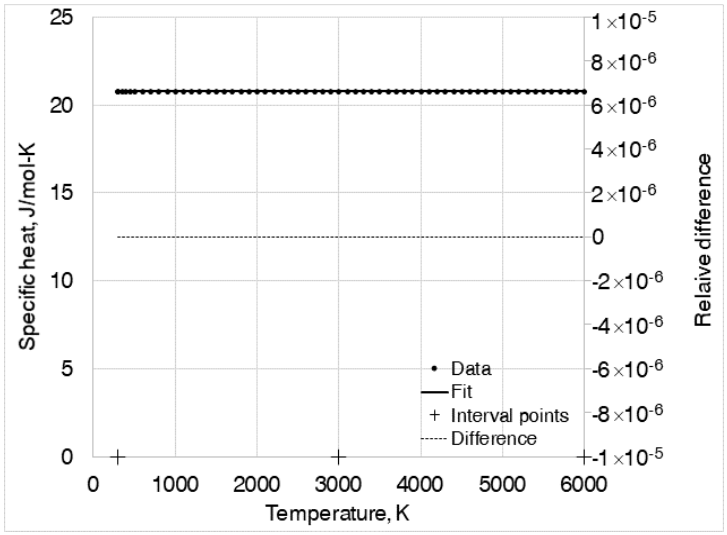


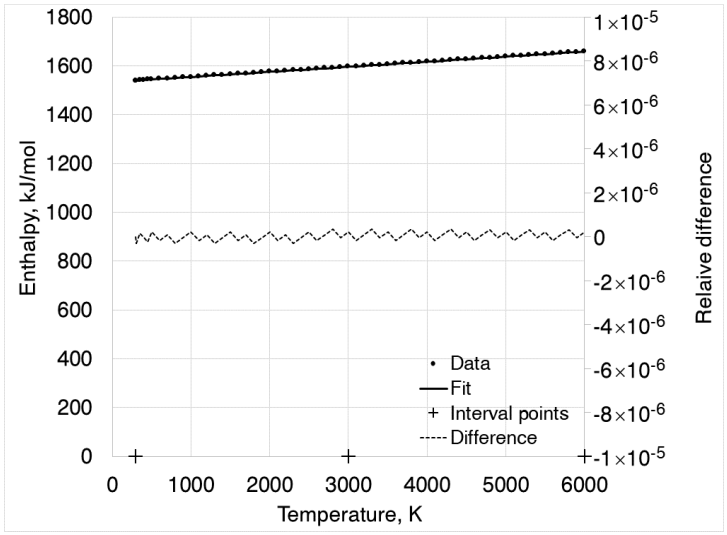


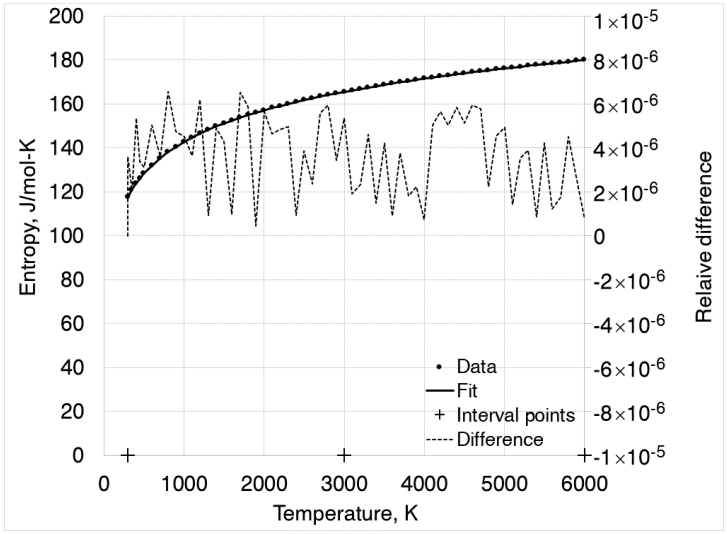


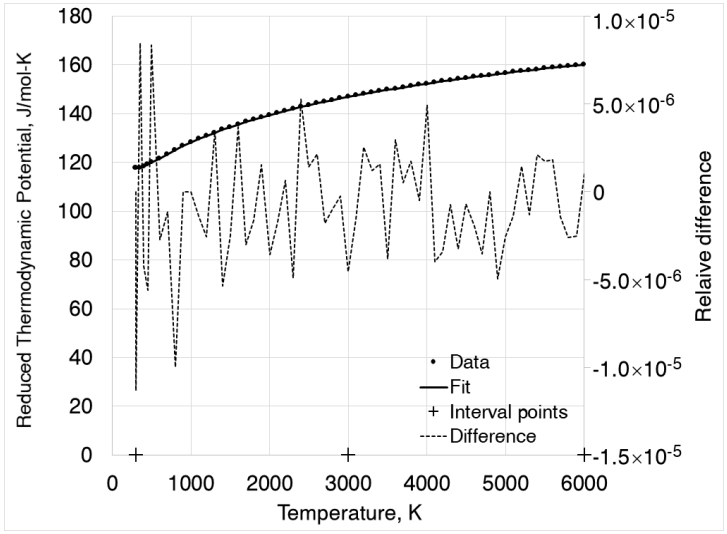
Deuterium, Ion (D^+) – D1+(g)

MW = 294 g/mol – $\Delta_f H^o$(298.15 K) = 1540.324 kJ/mol – S^o(298.15 K) = 117.583 J/mol-K – p^o = 0.1 MPa



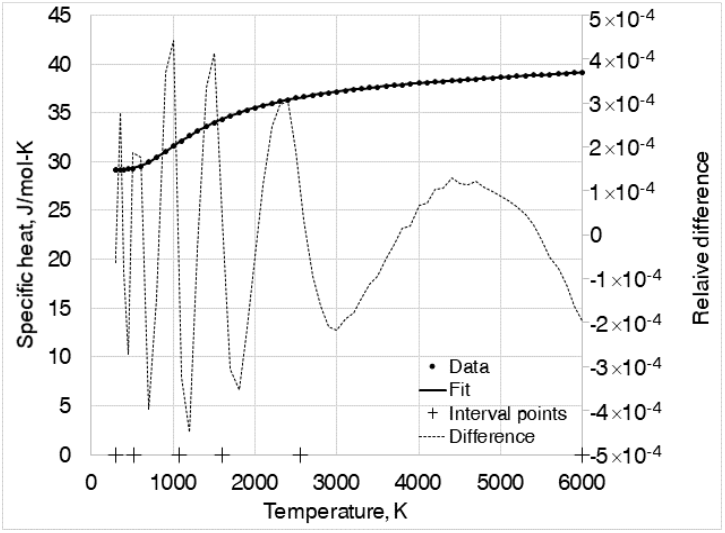


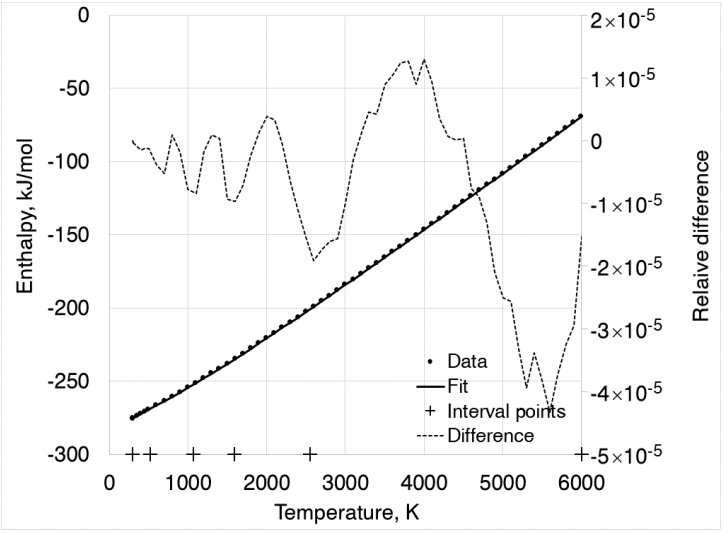


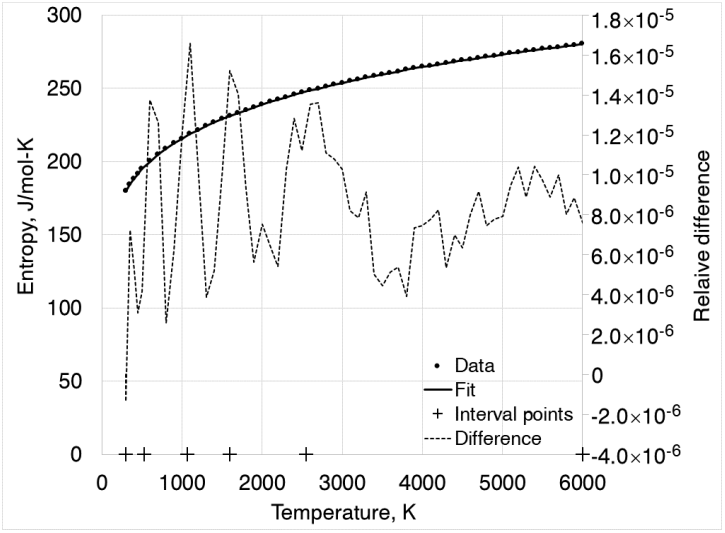


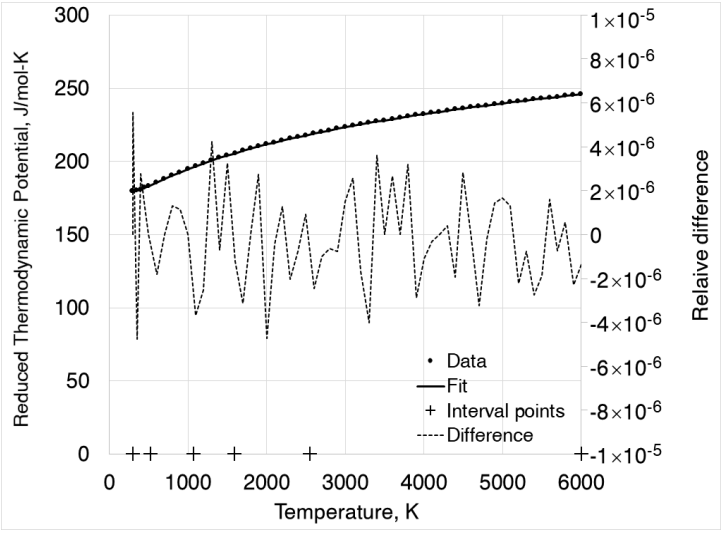
Hydrofluoric Acid-D (DF) – D1F1(g)

MW = 313 g/mol – $\Delta_f H^o$(298.15 K) = -275.516 kJ/mol – S^o(298.15 K) = 179.702 J/mol-K – p^o = 0.1 MPa



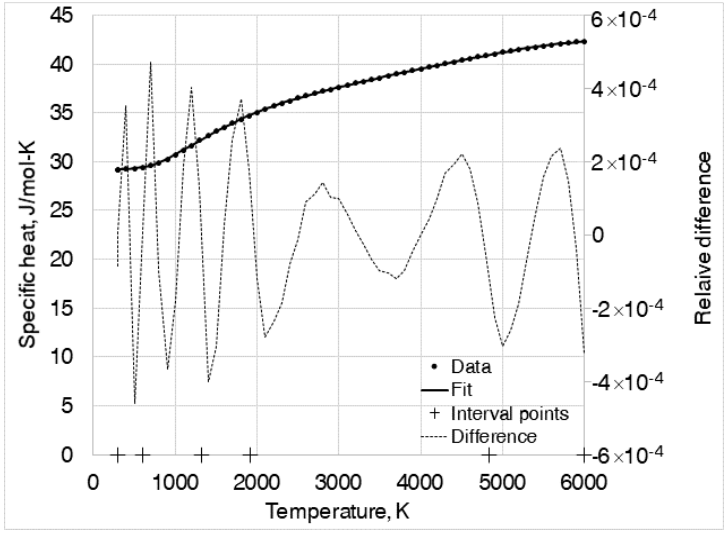


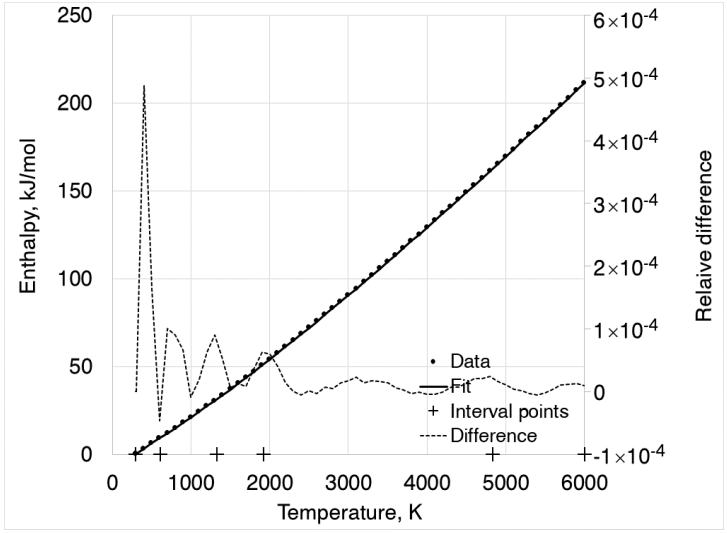


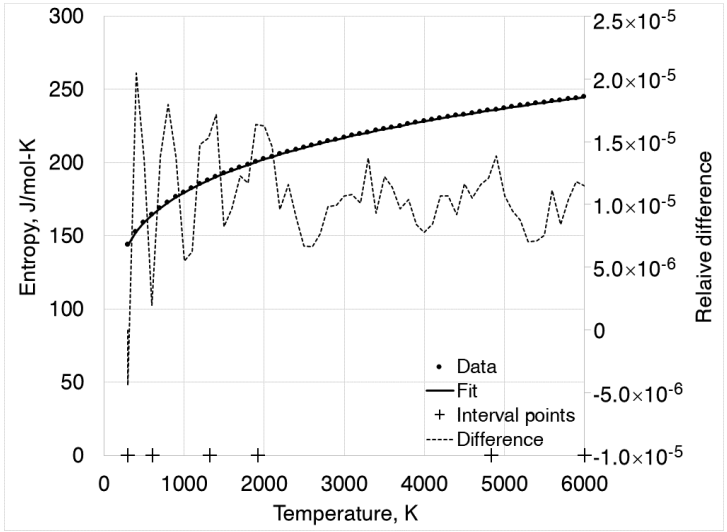


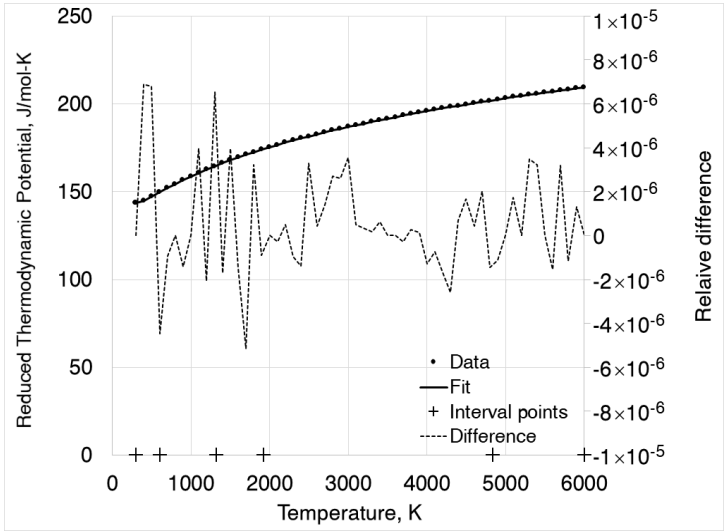
Hydrogen-D_1 (HD) – D1H1(g)

MW = 295 g/mol – $\Delta_f H^o$(298.15 K) = .32 kJ/mol – S^o(298.15 K) = 143.803 J/mol-K – p^o = 0.1 MPa



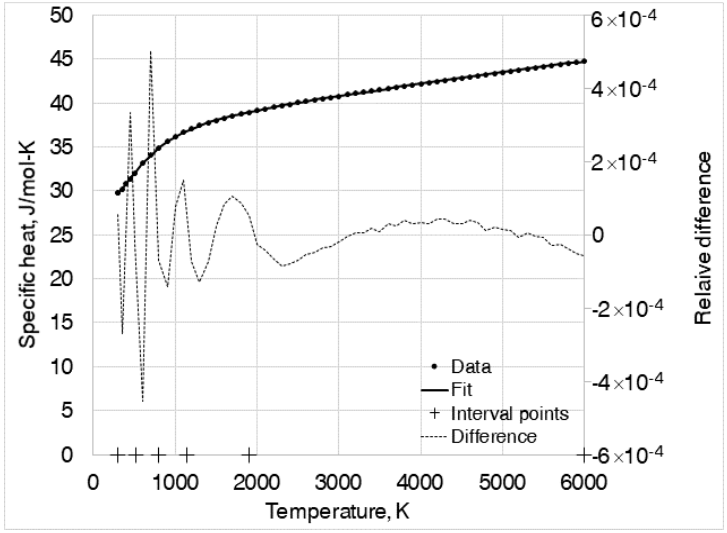


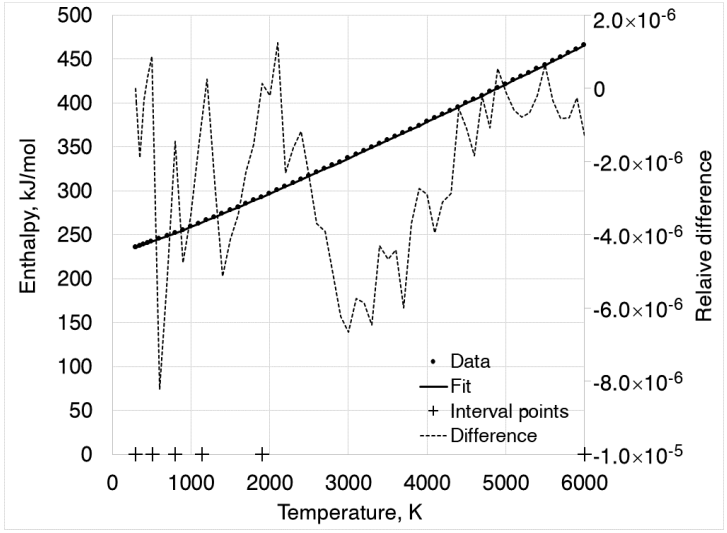


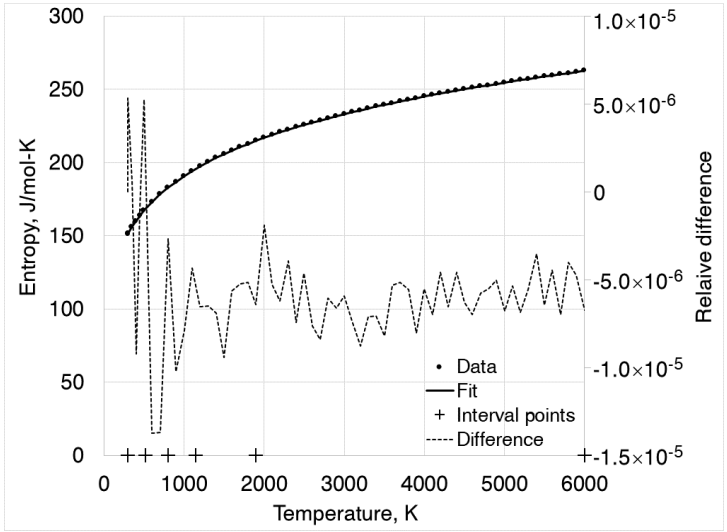


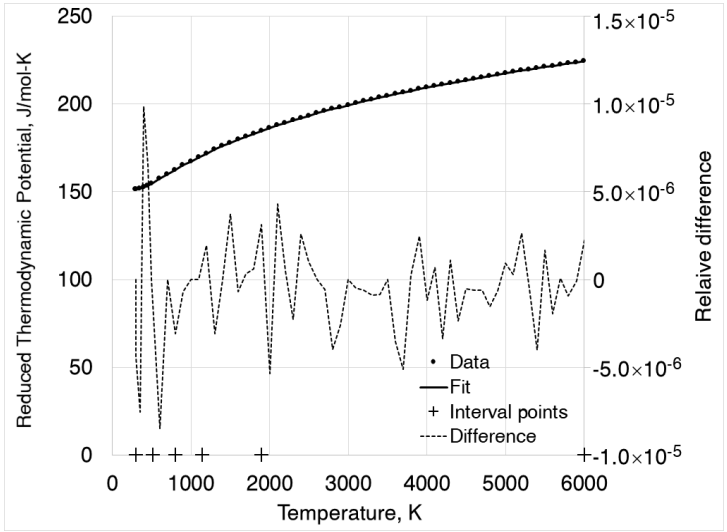
Hydrogen$^{-}D_1$, Ion (HD^{-}) – D1H1-(g)

MW = 295 g/mol – $\Delta_f H^o$(298.15 K) = 235.484 kJ/mol – S^o(298.15 K) = 151.206 J/mol-K – p^o = 0.1 MPa



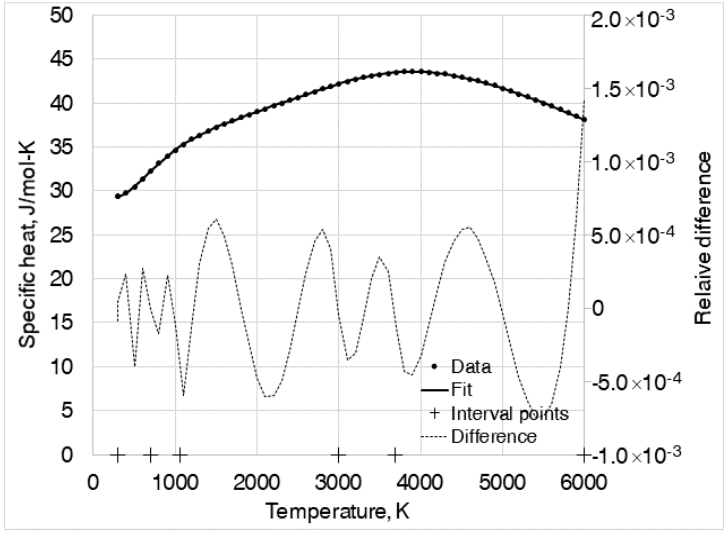


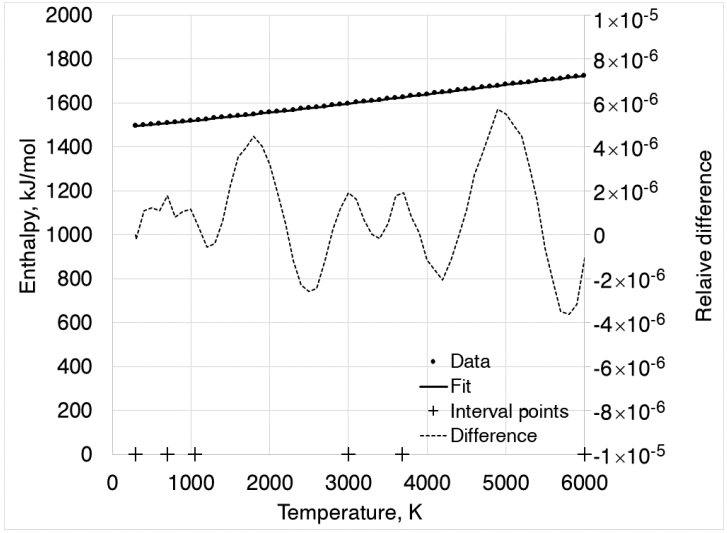


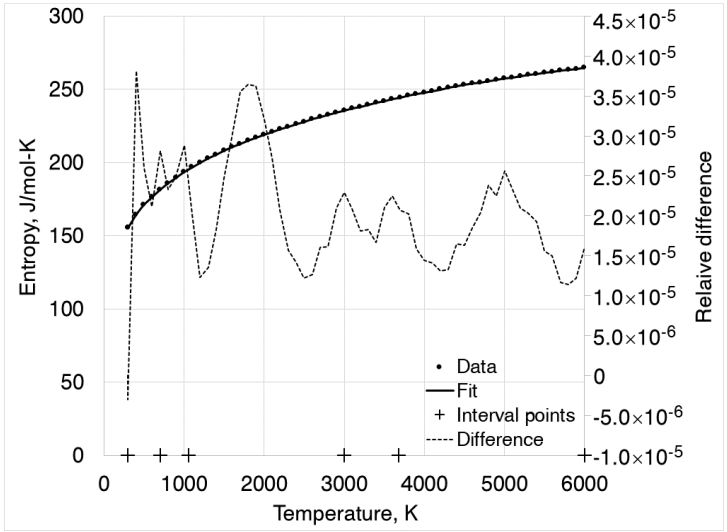


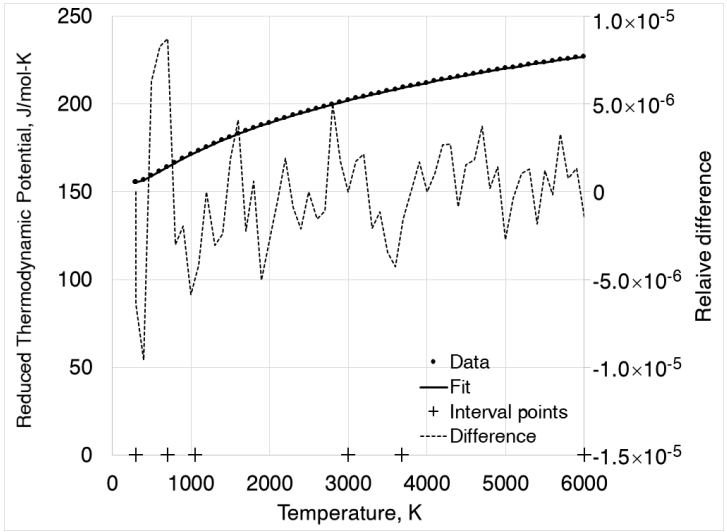
Hydrogen$^{-}D_1$, Ion (HD^{+}) – D1H1+(g)

MW = 295 g/mol – $\Delta_f H^o$(298.15 K) = 1496.793 kJ/mol – S^o(298.15 K) = 155.552 J/mol-K – p^o = 0.1 MPa



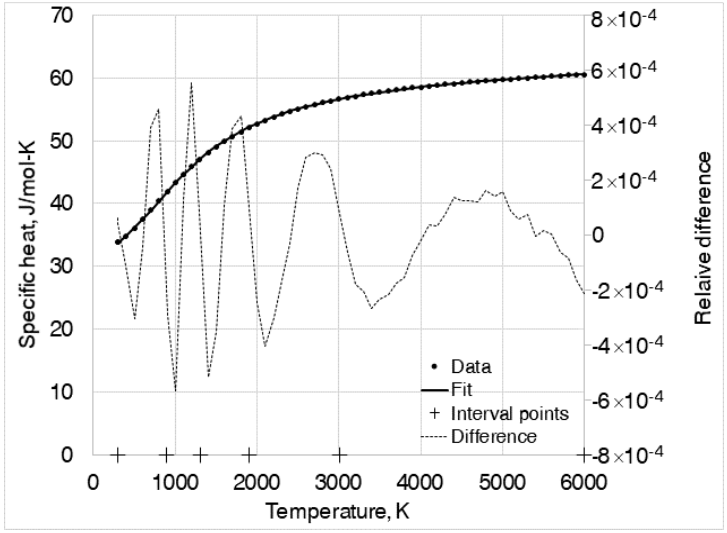


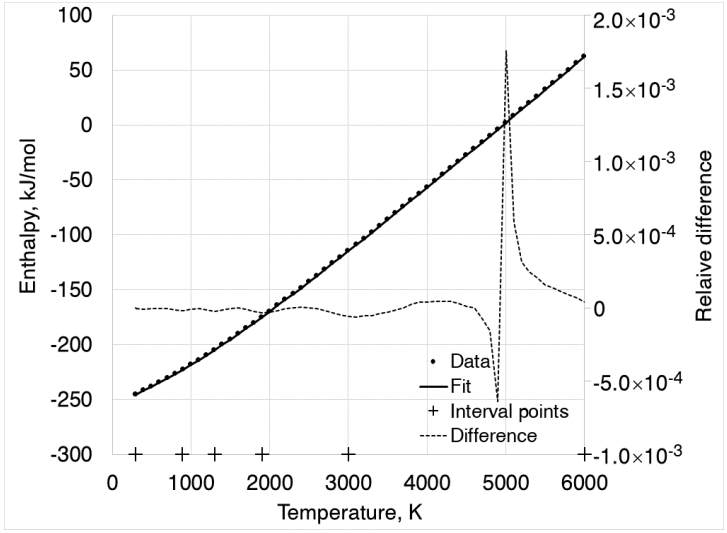


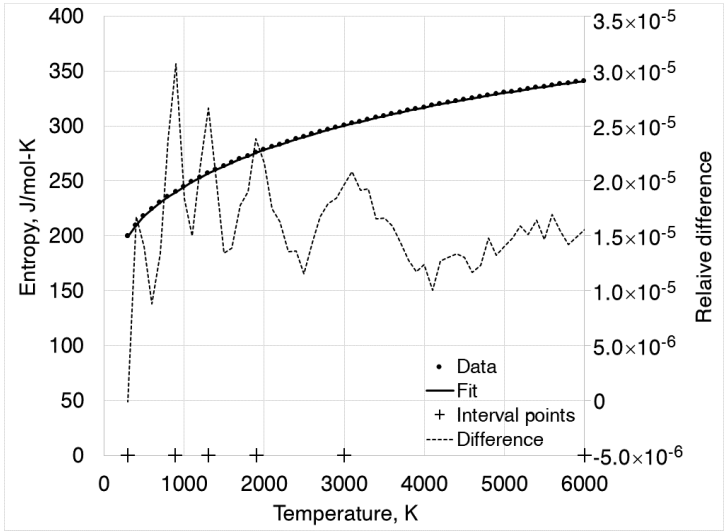


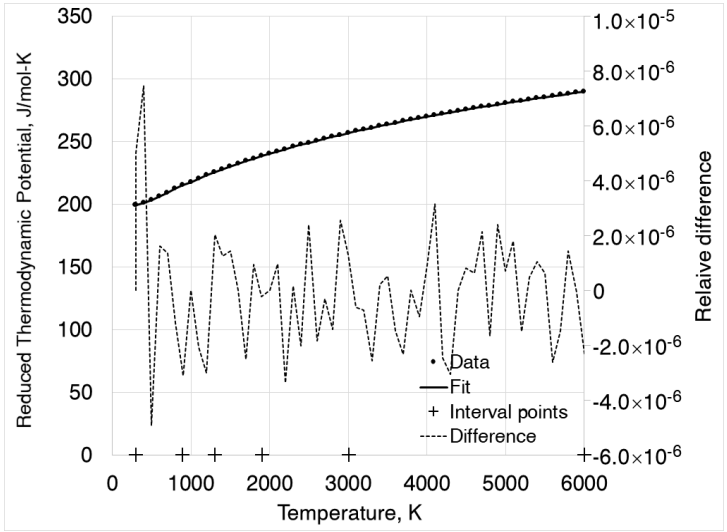
Water-D_1 (HDO) – D1H1O1(g)

MW = 311 g/mol – $\Delta_f H^o$(298.15 K) = -245.371 kJ/mol – S^o(298.15 K) = 199.511 J/mol-K – p^o = 0.1 MPa



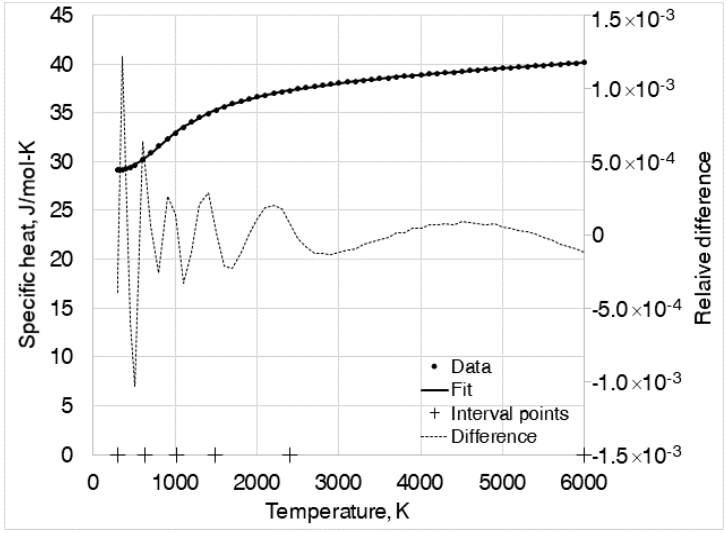


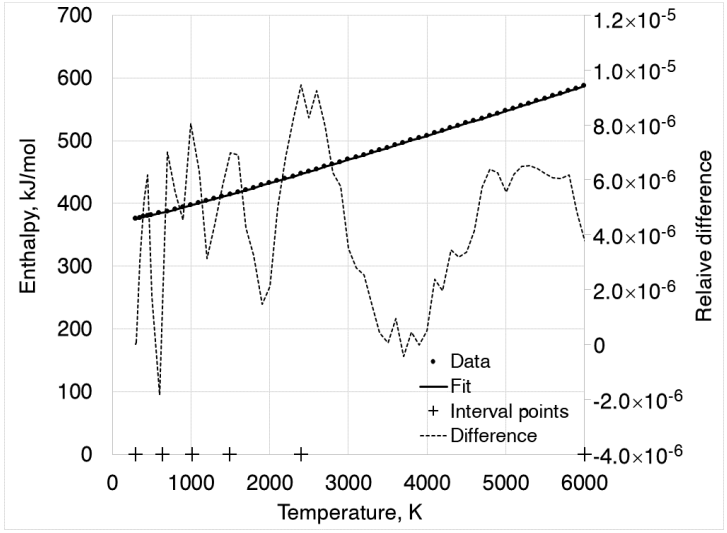


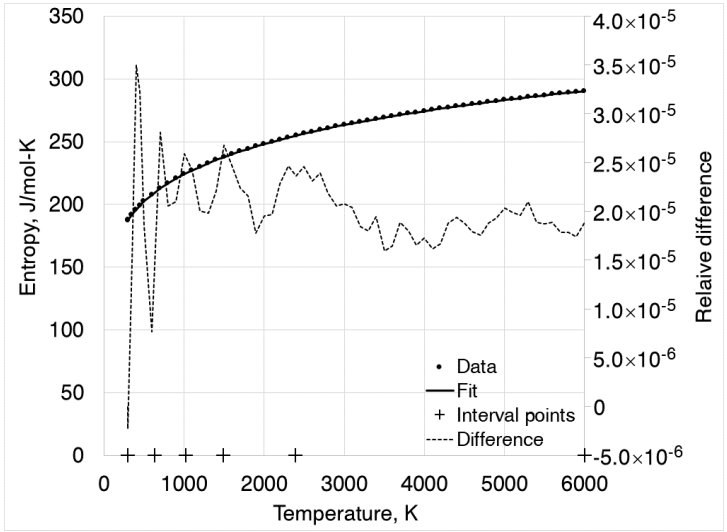


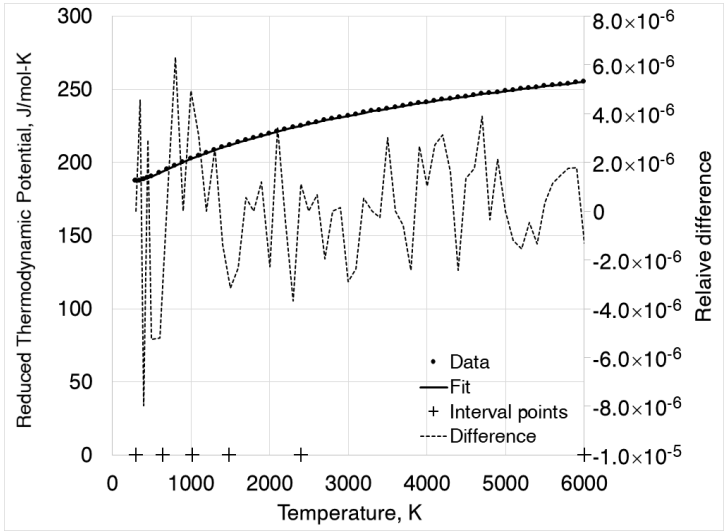
Imidogen-D_1 (ND) – D1N1(g)

MW = 308 g/mol – $\Delta_f H^o$(298.15 K) = 375.305 kJ/mol – S^o(298.15 K) = 187.234 J/mol-K – p^o = 0.1 MPa



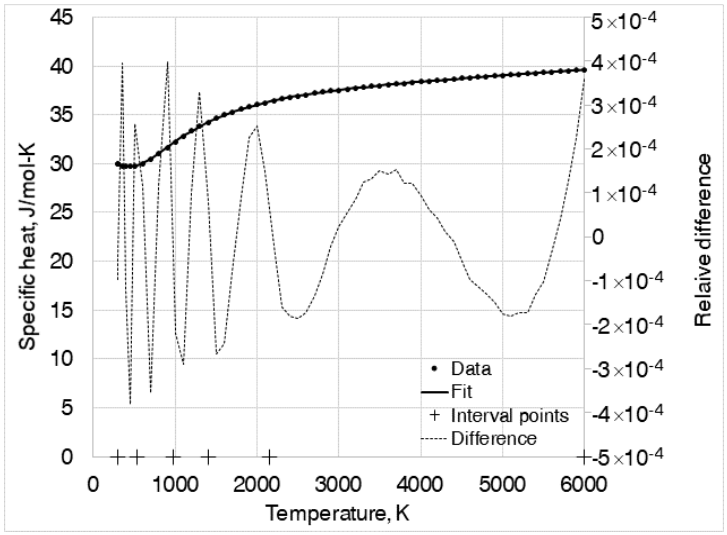


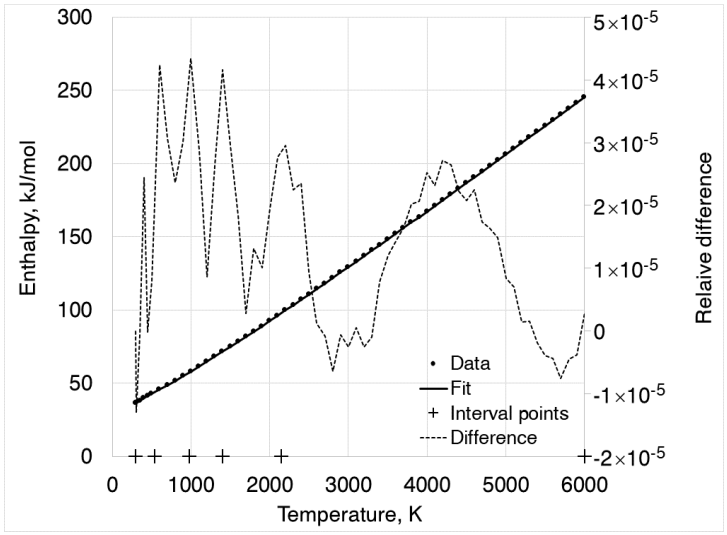


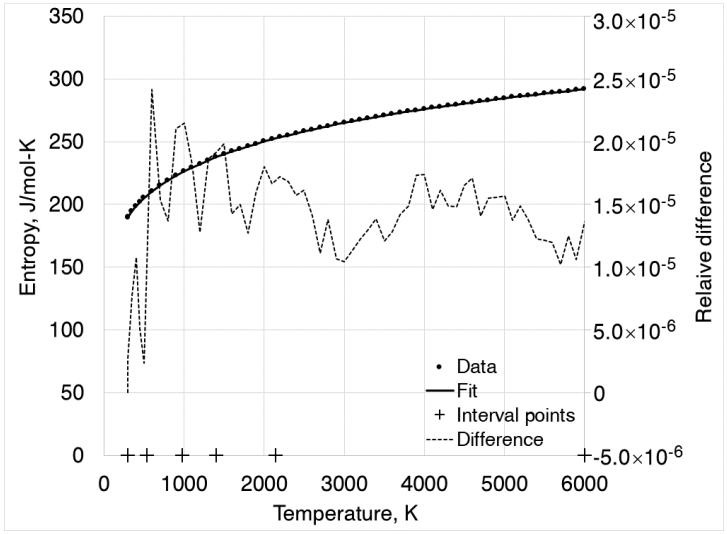


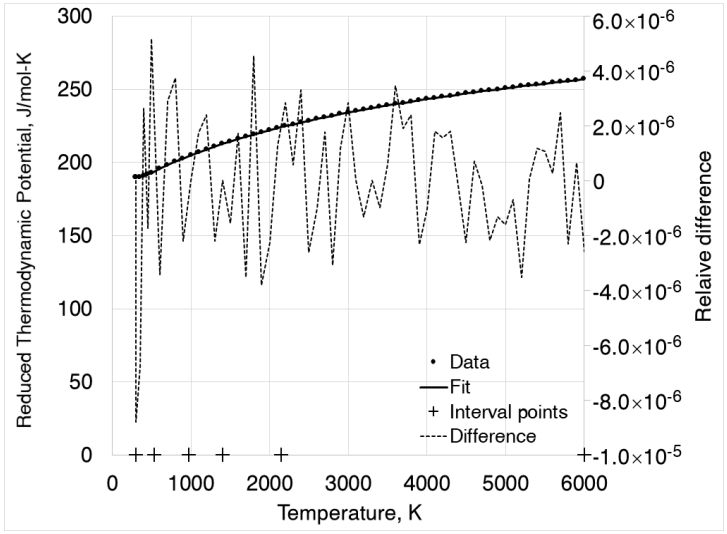
Hydroxyl$^{-}D_1$ (OD) – D1O1(g)

MW = 310 g/mol – $\Delta_f H^o$(298.15 K) = 36.602 kJ/mol – S^o(298.15 K) = 189.651 J/mol-K – p^o = 0.1 MPa



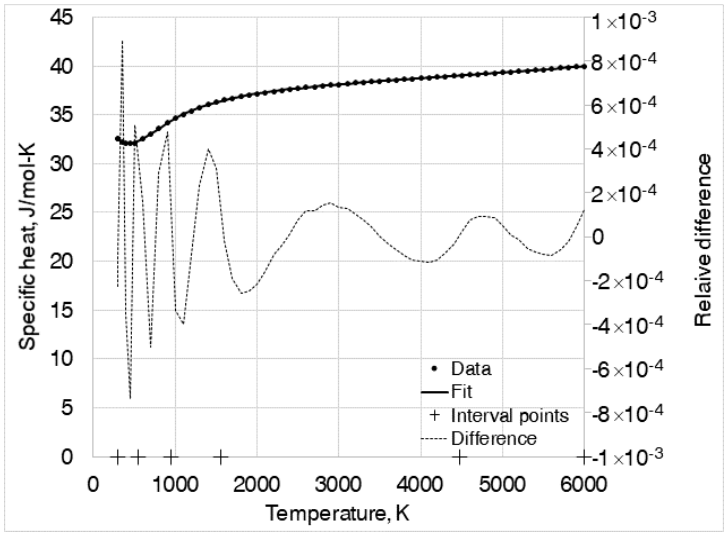


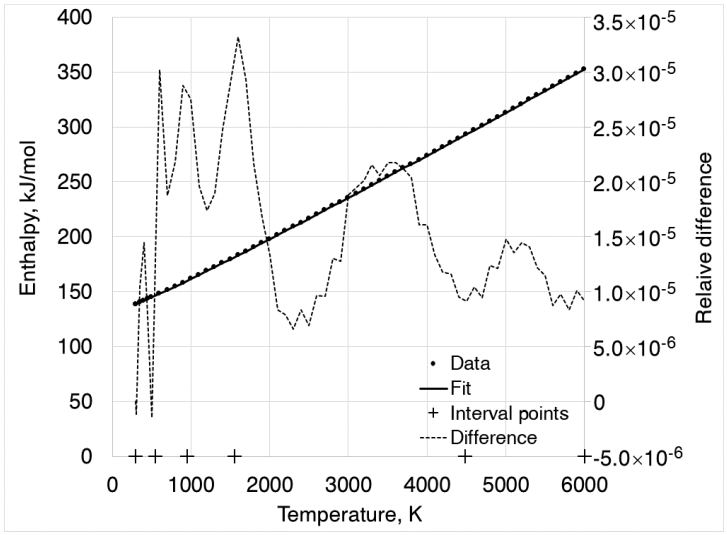


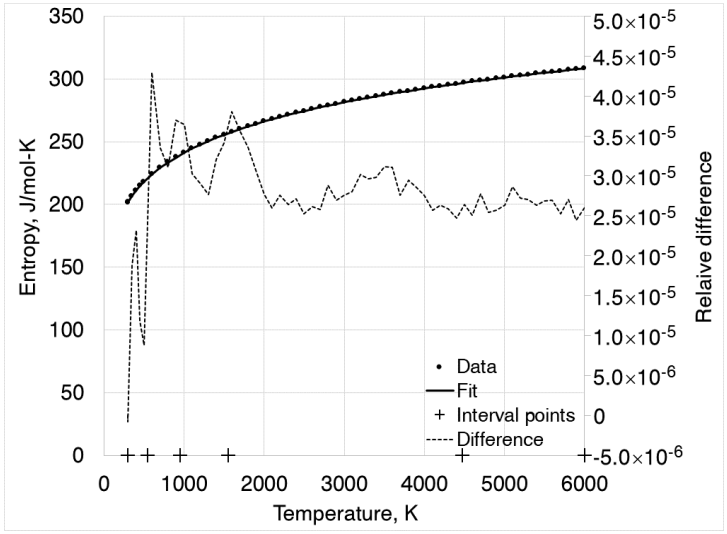


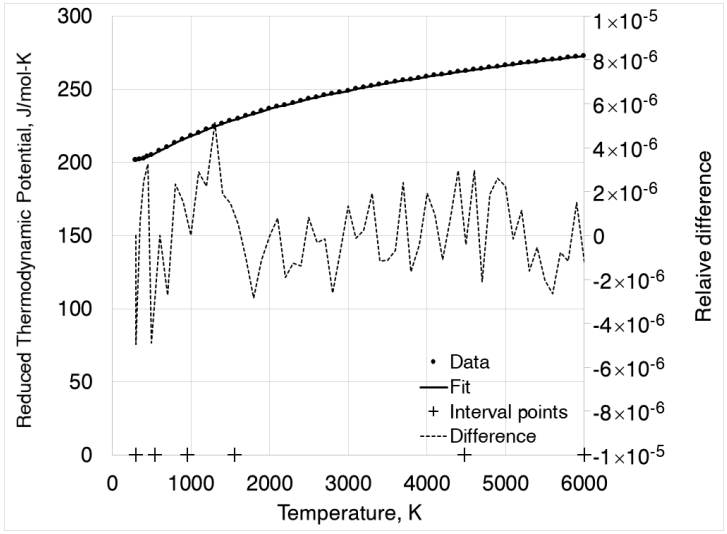
Mercapto-D_1 (SD) – D1S1(g)

MW = 326 g/mol – $\Delta_f H^o$(298.15 K) = 138.49 kJ/mol – S^o(298.15 K) = 201.484 J/mol-K – p^o = 0.1 MPa



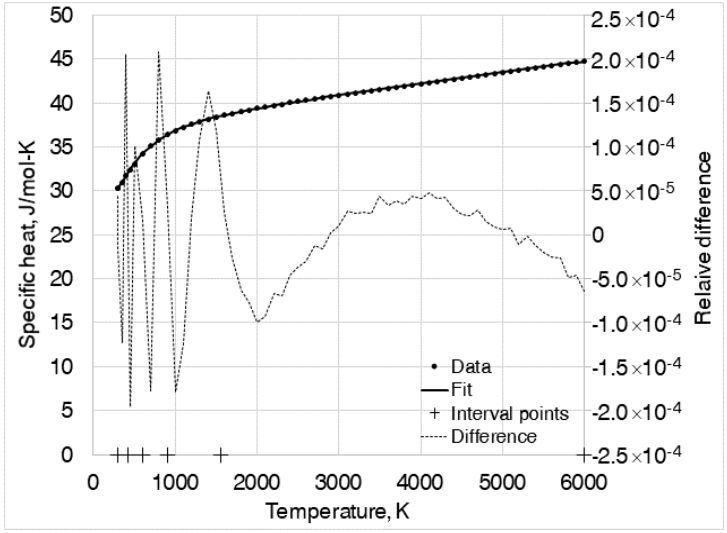


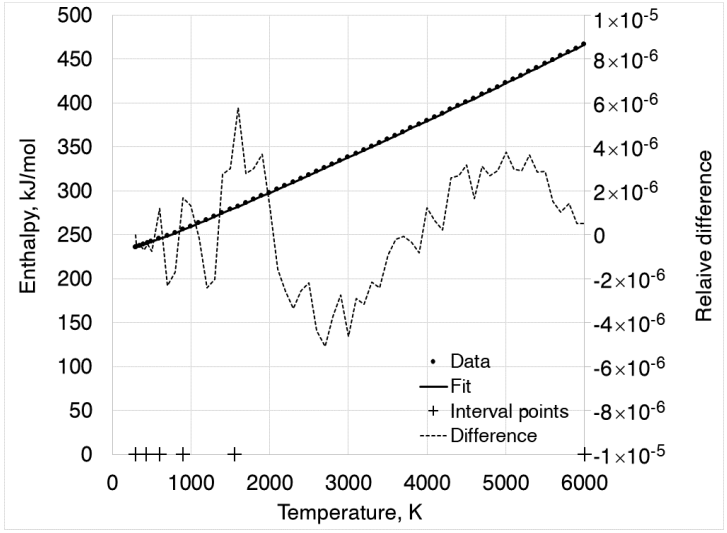


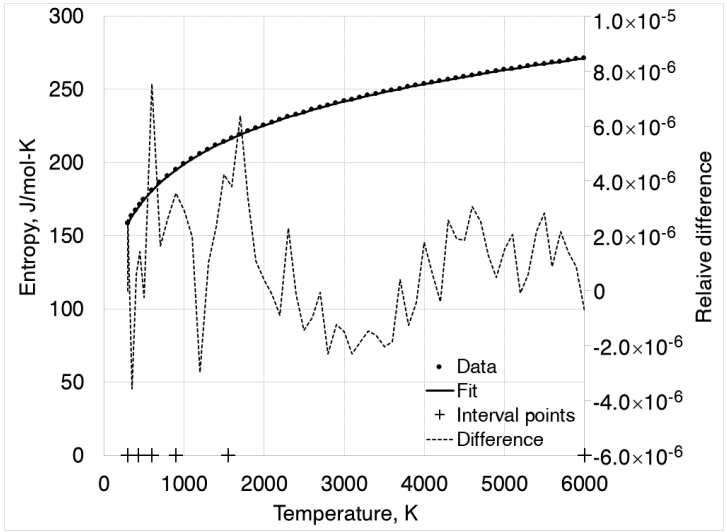


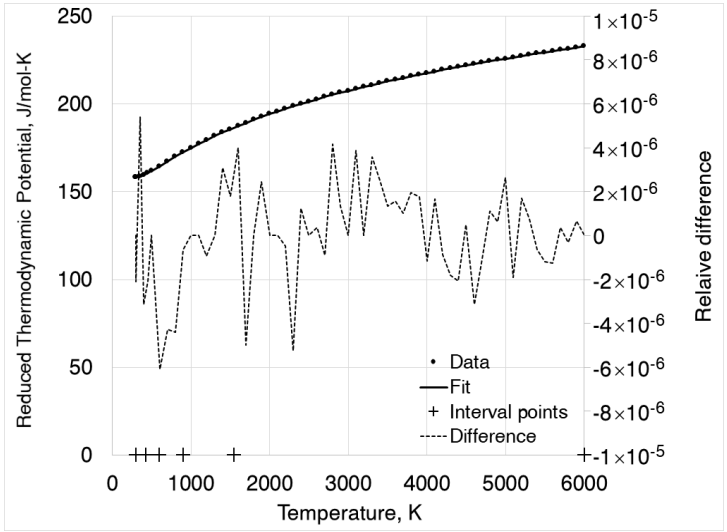
Deuterium, Ion (D_2^-) – D2-(g)

MW = 588 g/mol – $\Delta_f H^o$(298.15 K) = 235.365 kJ/mol – S^o(298.15 K) = 158.259 J/mol-K – p^o = 0.1 MPa



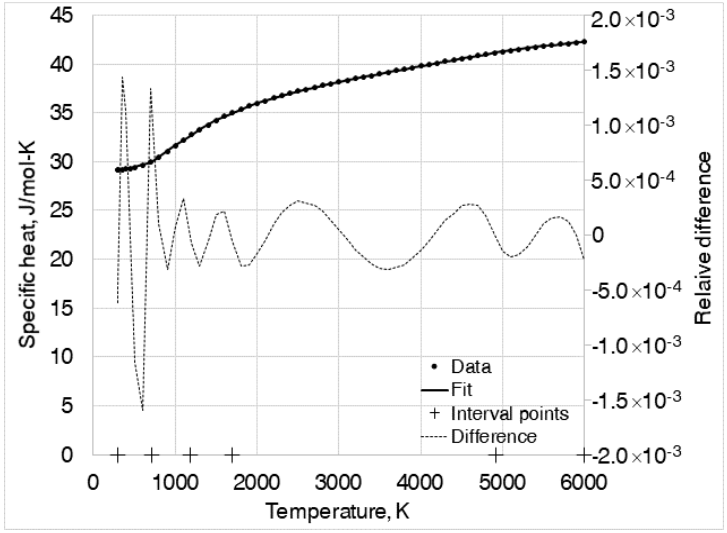


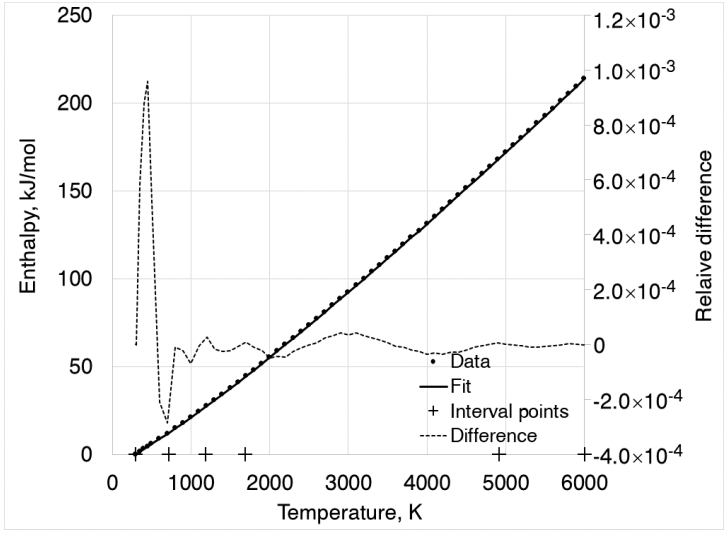


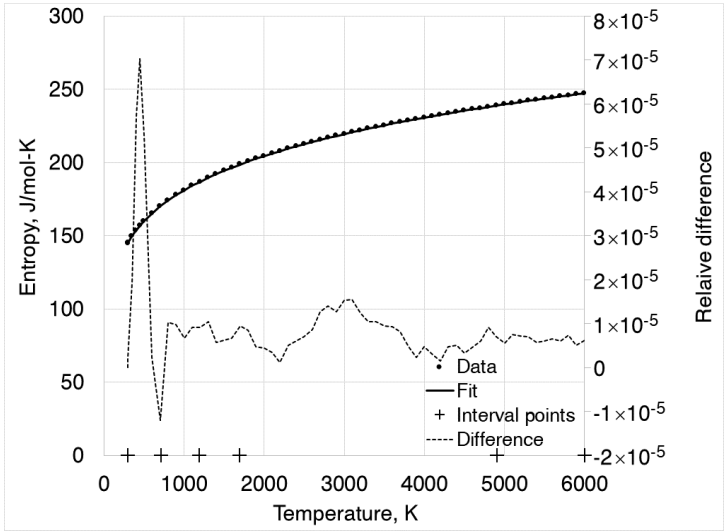


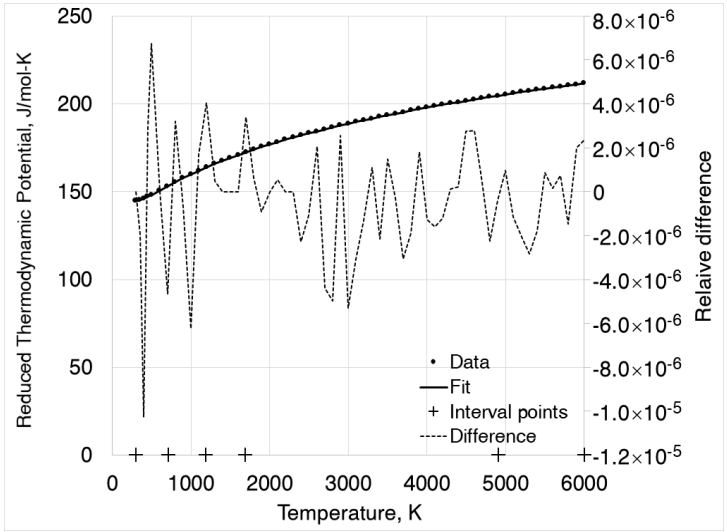
Deuterium (D_2) – D2(ref)

MW = 588 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 144.96 J/mol-K – p^o = 0.1 MPa



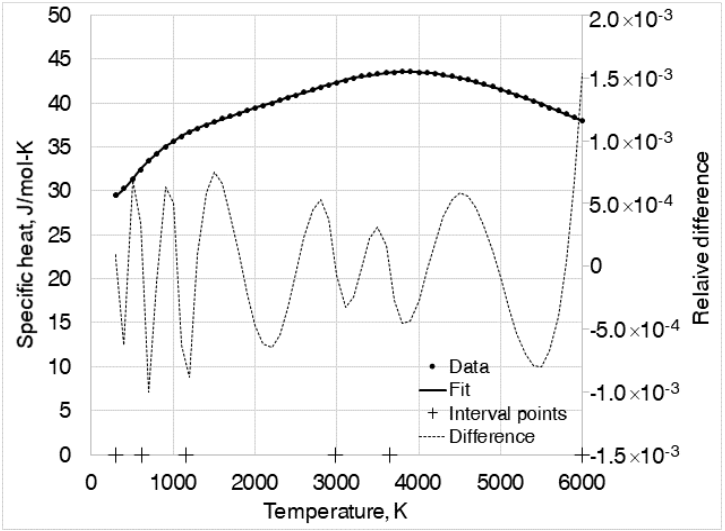


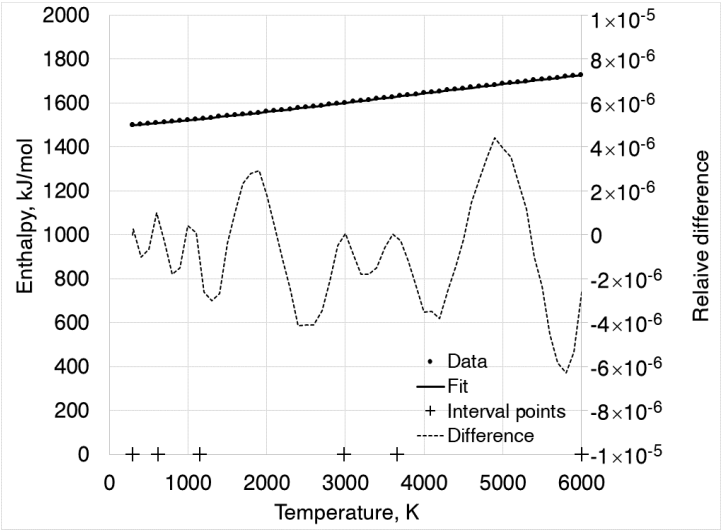


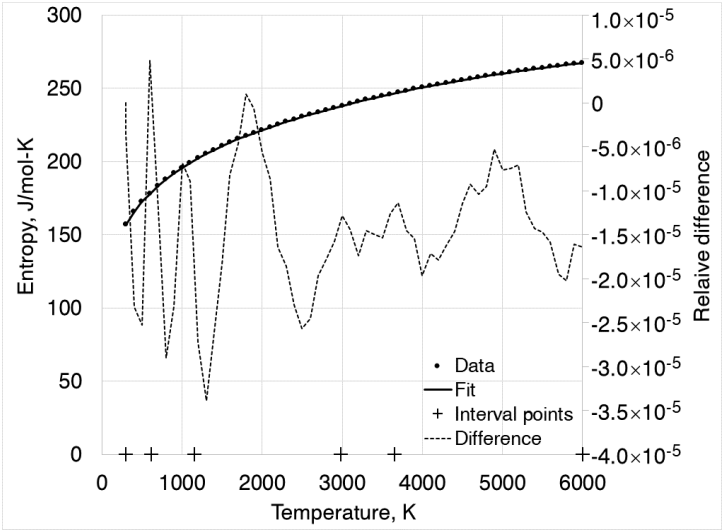


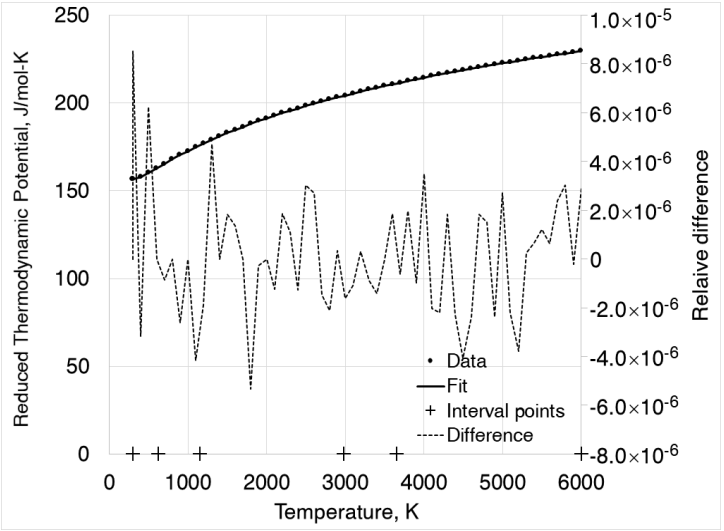
Deuterium, Ion (D_2^+) – D2+(g)

MW = 588 g/mol – $\Delta_f H^o$(298.15 K) = 1498.567 kJ/mol – S^o(298.15 K) = 156.735 J/mol-K – p^o = 0.1 MPa



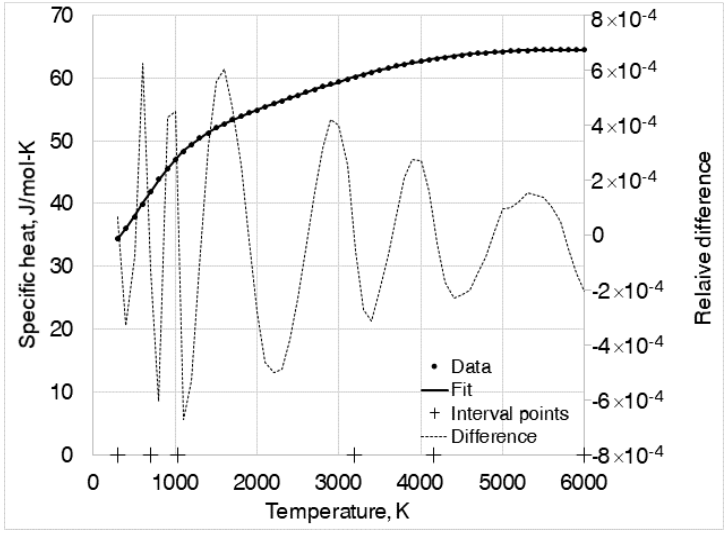


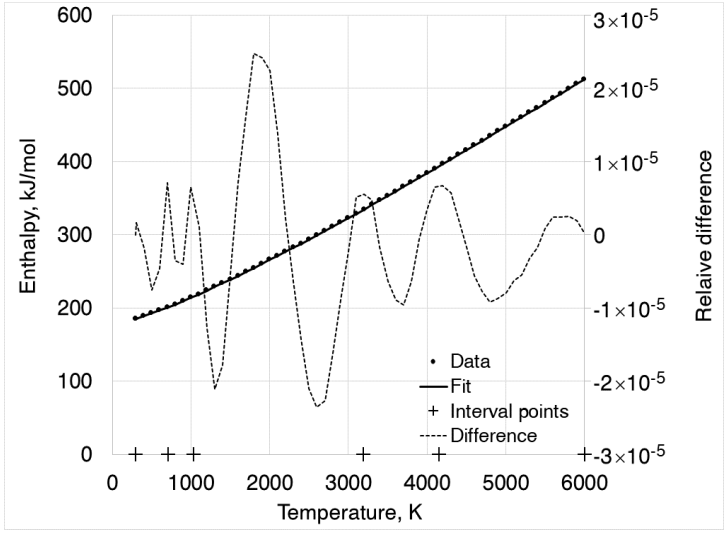


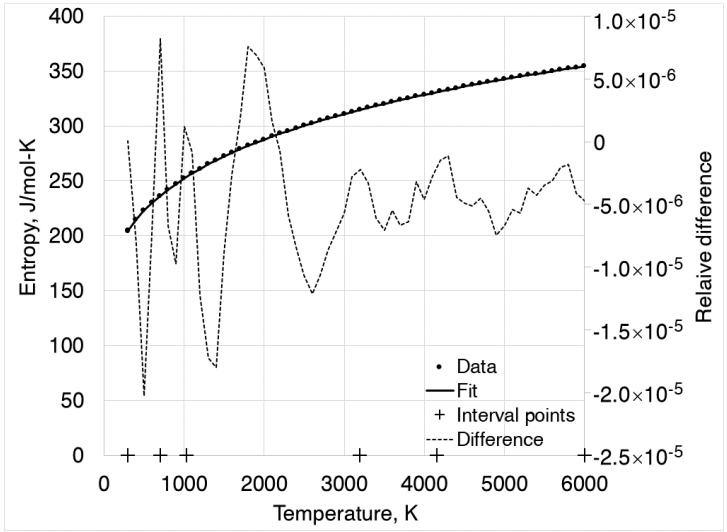


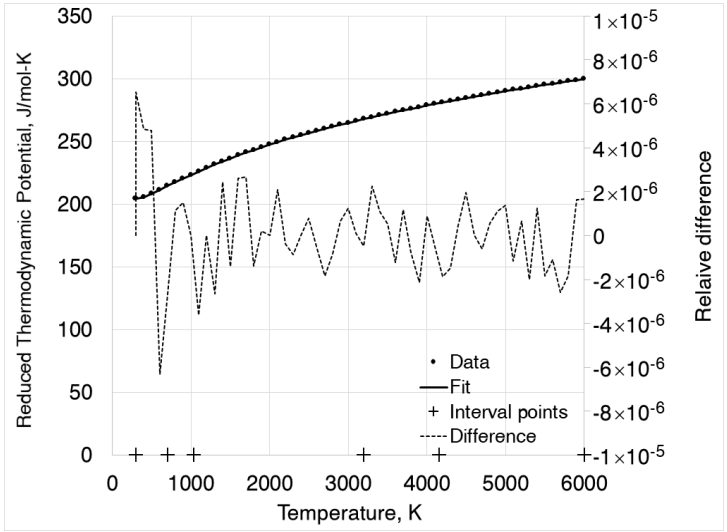
Amidogen-D_2 (ND_2) – D2N1(g)

MW = 602 g/mol – $\Delta_f H^o$(298.15 K) = 185.351 kJ/mol – S^o(298.15 K) = 204.291 J/mol-K – p^o = 0.1 MPa



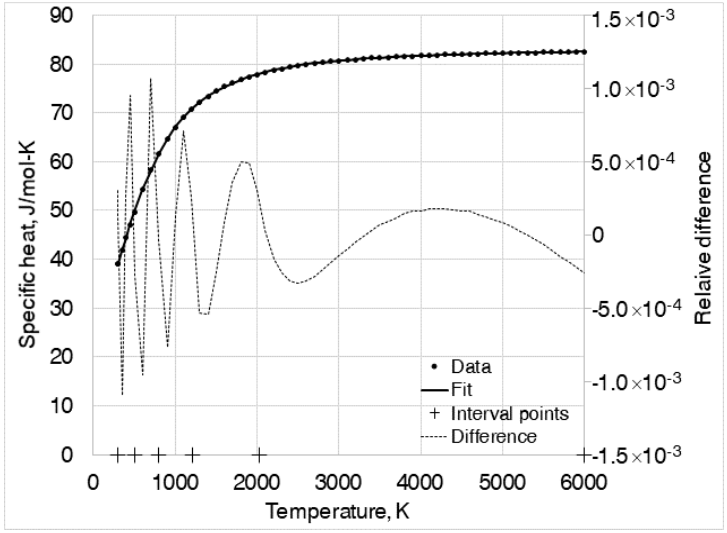


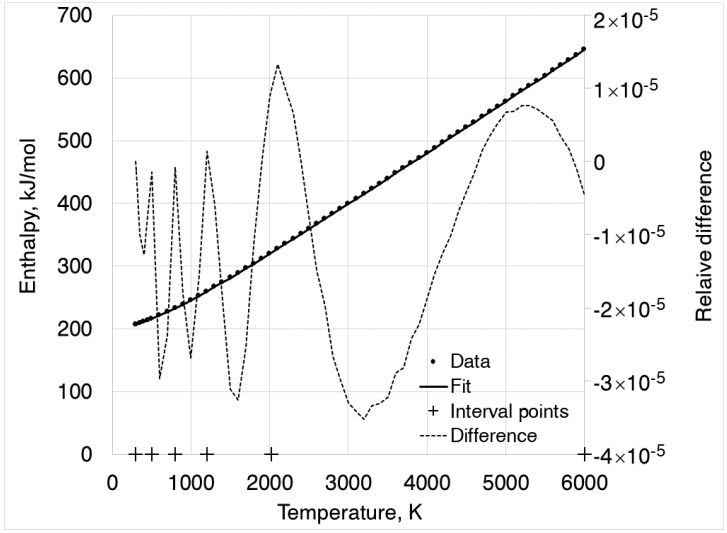


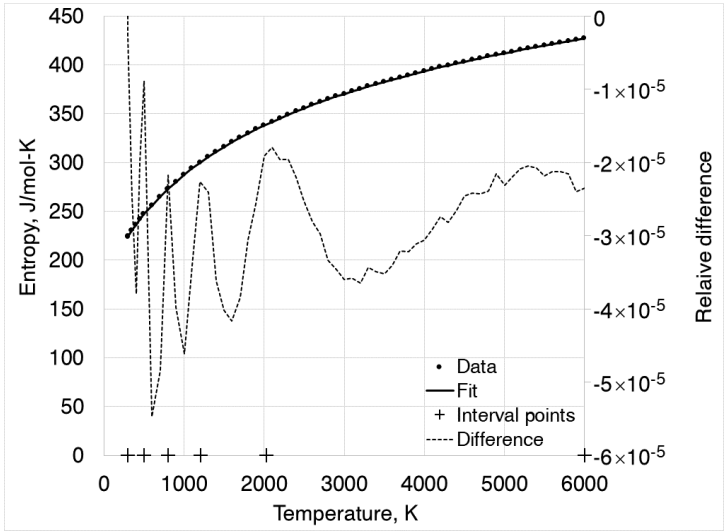


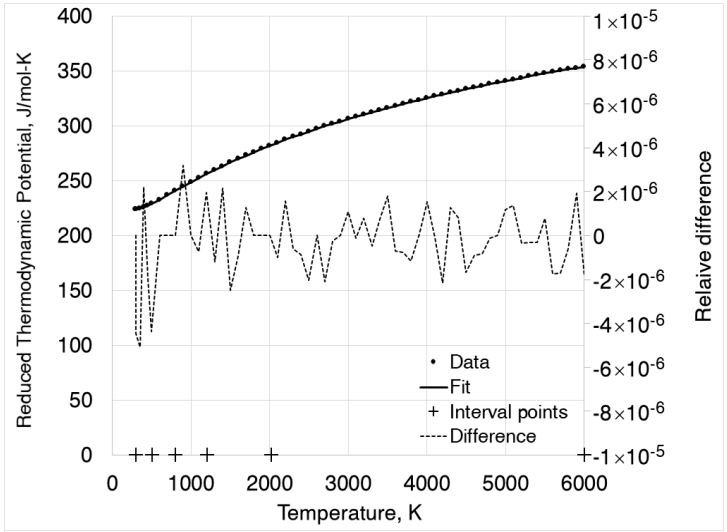
Diazene-D_2, Cis (DNND) – D2N2(g)

MW = 616 g/mol – $\Delta_f H^o$(298.15 K) = 207.108 kJ/mol – S^o(298.15 K) = 224.092 J/mol-K – p^o = 0.1 MPa



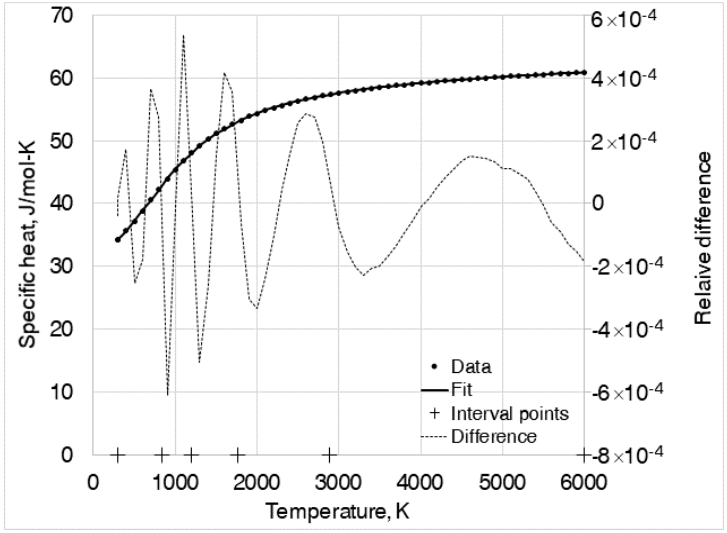


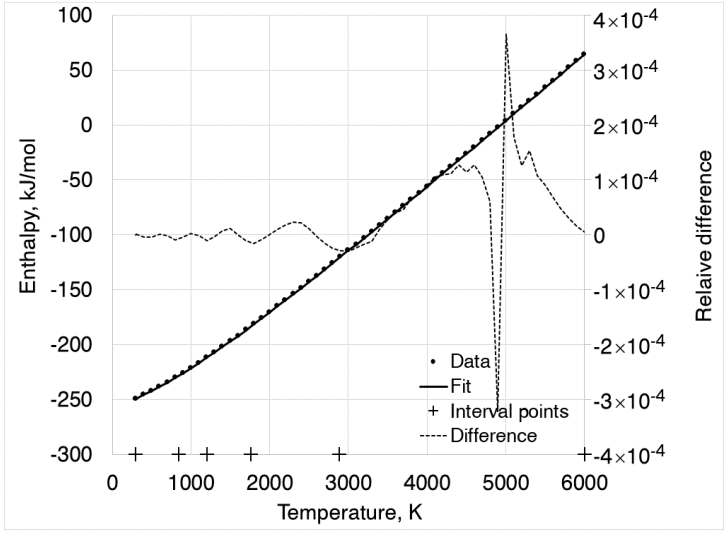


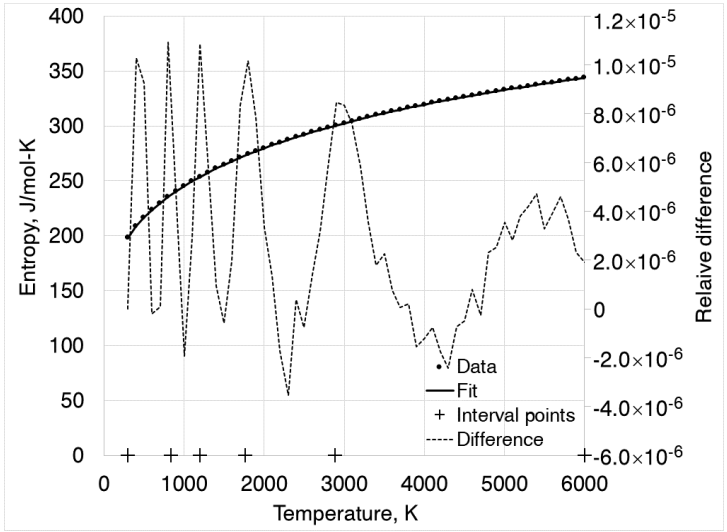


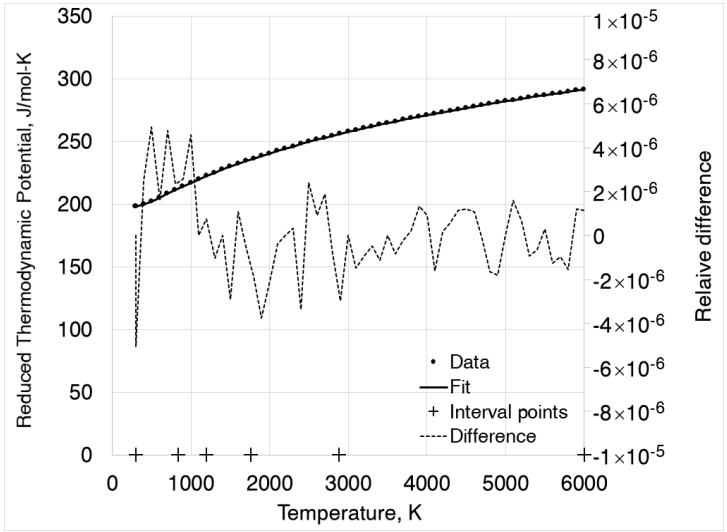
Water-D_2 (D_2O) – D2O1(g)

MW = 604 g/mol – $\Delta_f H^o$(298.15 K) = -249.203 kJ/mol – S^o(298.15 K) = 198.339 J/mol-K – p^o = 0.1 MPa



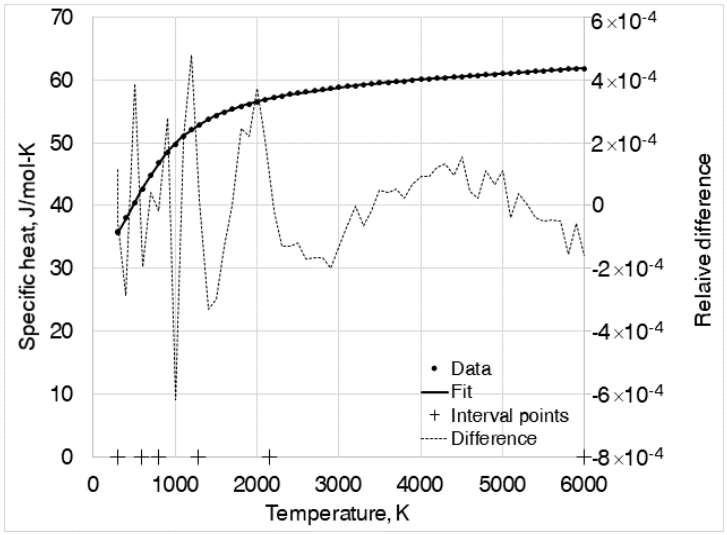


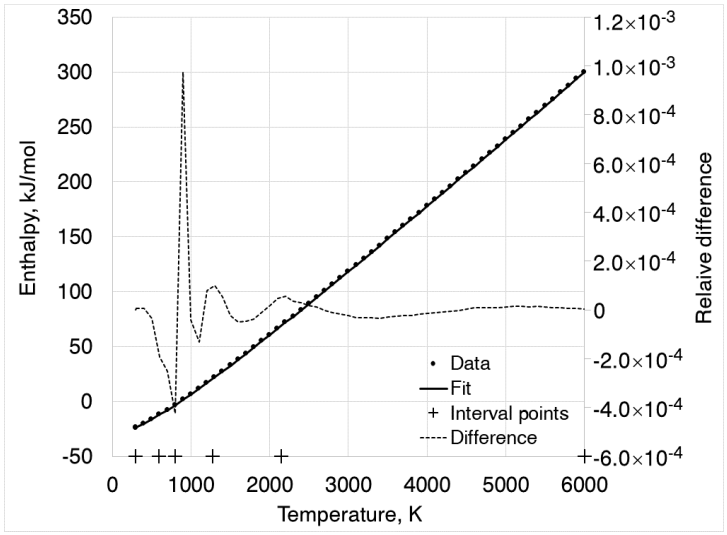


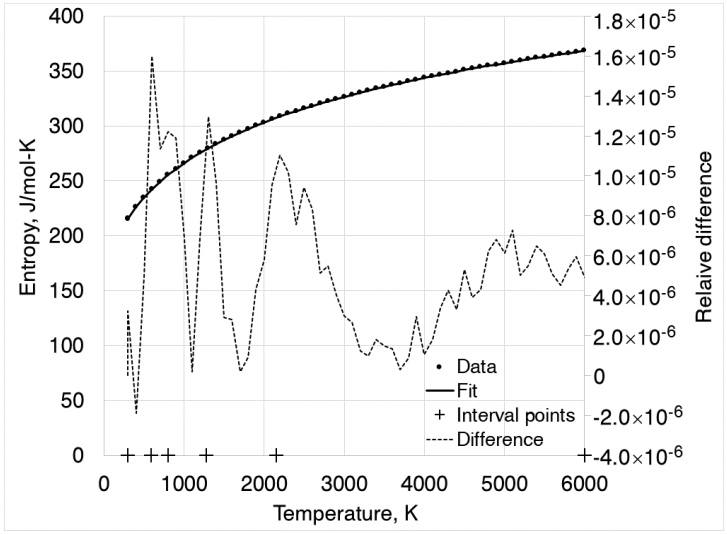


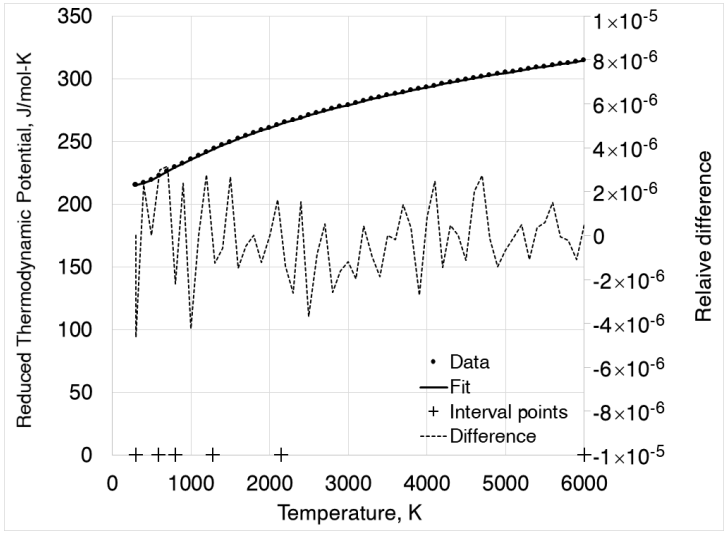
Hydrogen Sulfide-D_2 (D_2S) – D2S1(g)

MW = 620 g/mol – $\Delta_f H^o$(298.15 K) = -23.891 kJ/mol – S^o(298.15 K) = 215.284 J/mol-K – p^o = 0.1 MPa



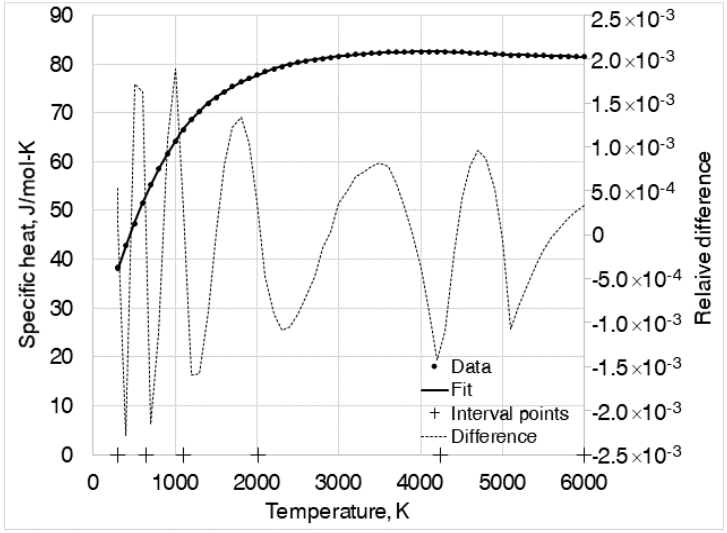


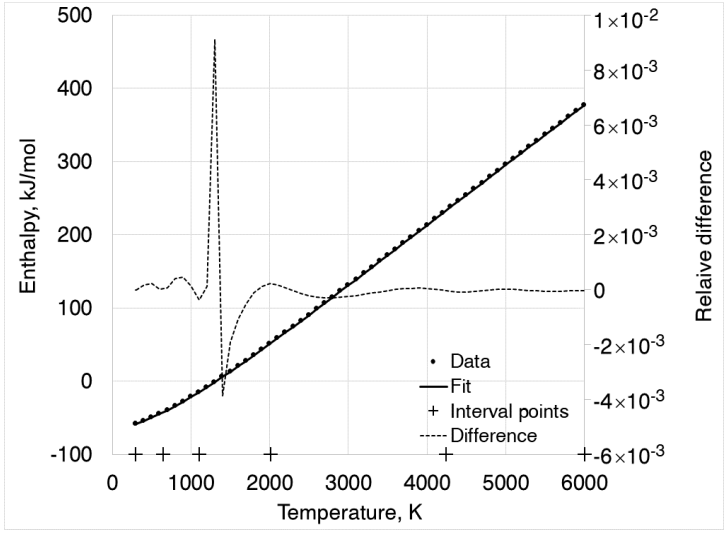


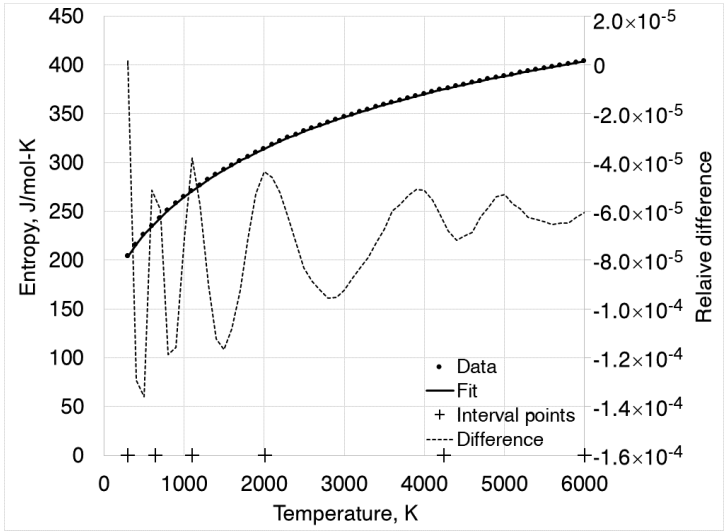


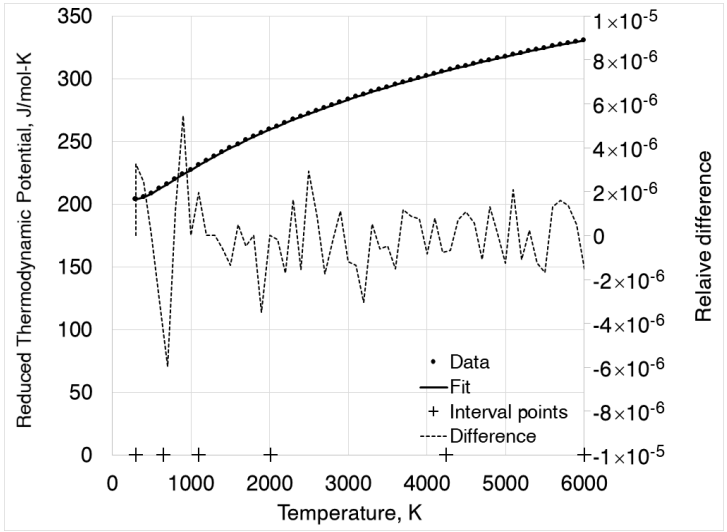
Ammonia-D_3 (ND_3) – D3N1(g)

MW = 896 g/mol – $\Delta_f H^o$(298.15 K) = -58.576 kJ/mol – S^o(298.15 K) = 203.931 J/mol-K – p^o = 0.1 MPa



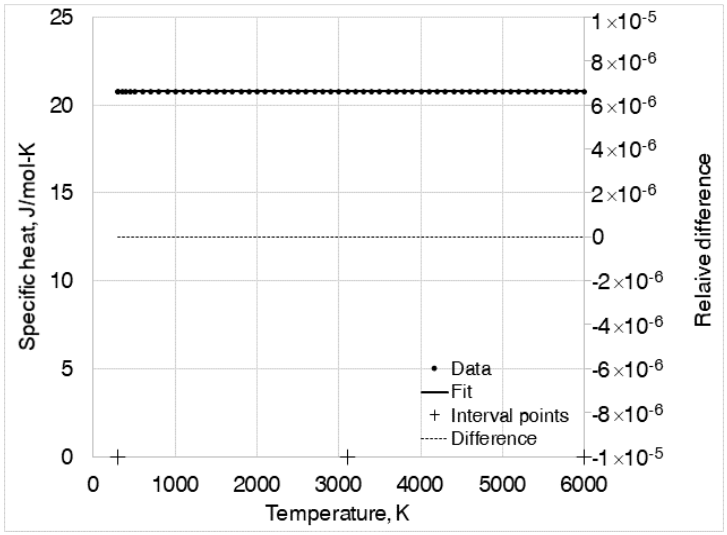


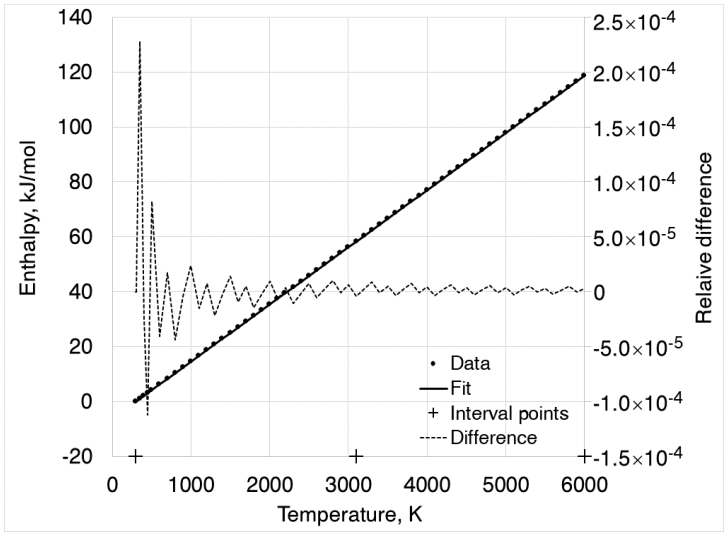


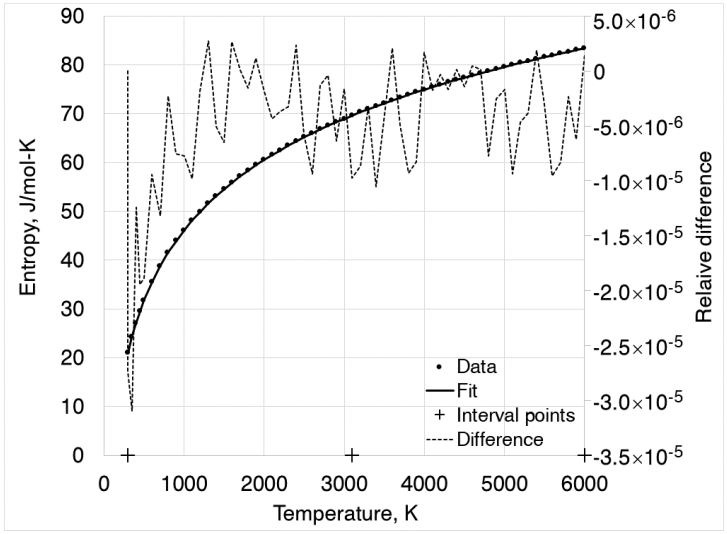


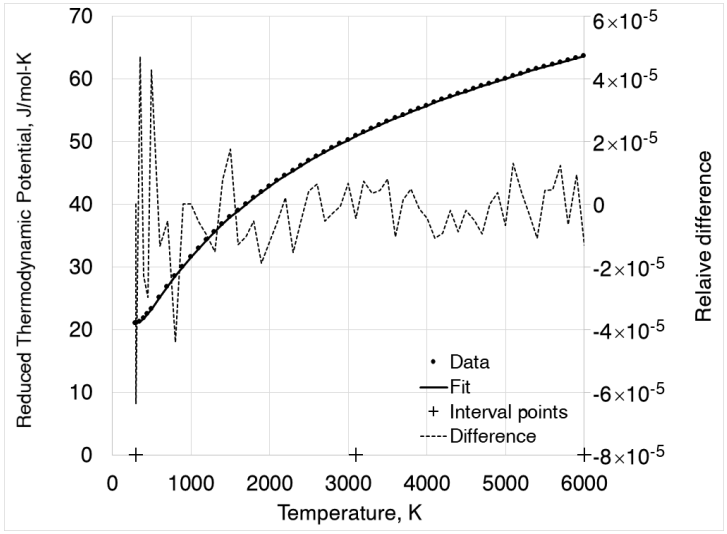
Electron Gas (e^-) – e1-(ref)

MW = .00054858 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 20.979 J/mol-K – p^o = 0.1 MPa



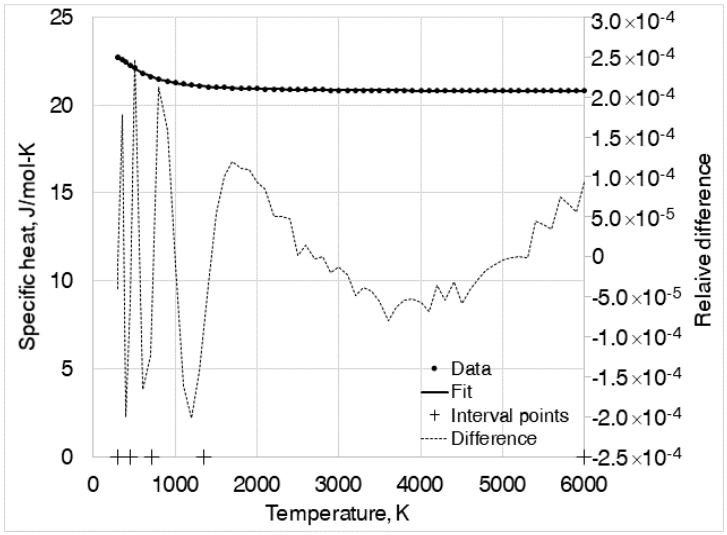


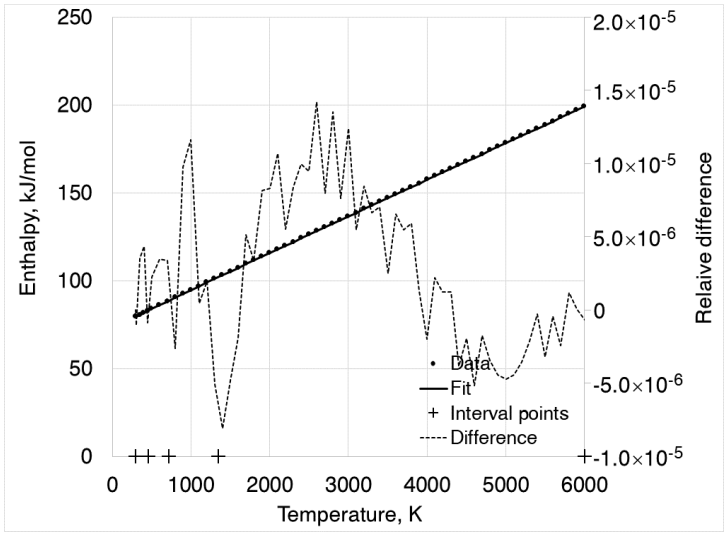


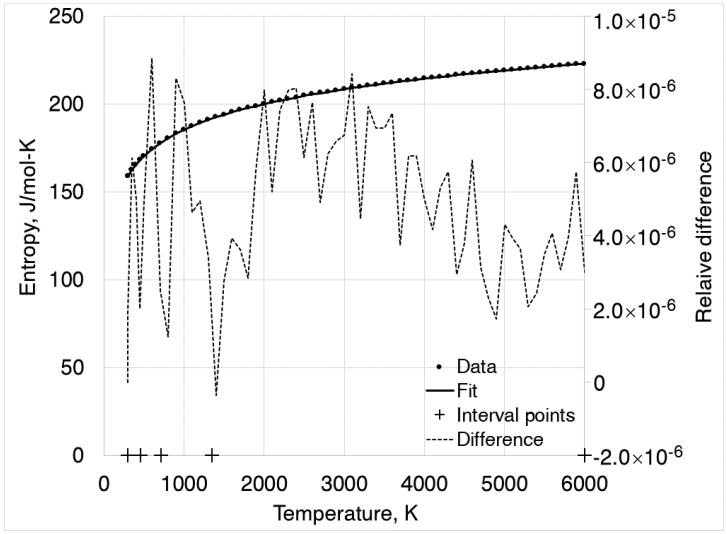


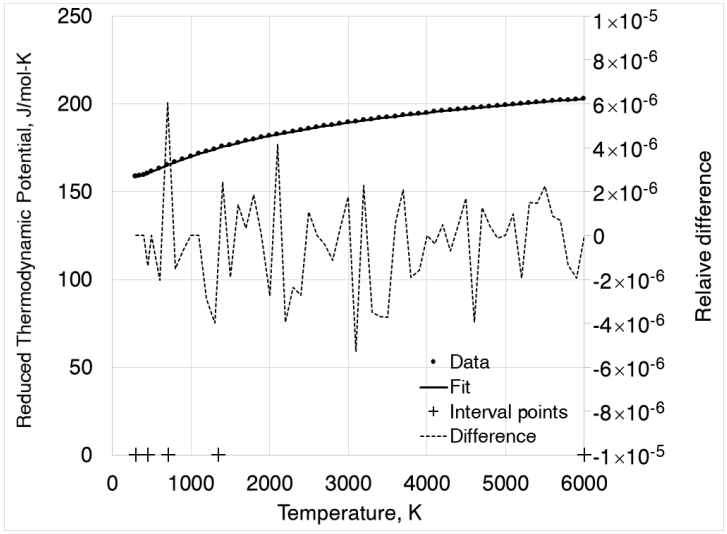
Fluorine (F) – F1(g)

MW = 18.9984 g/mol – $\Delta_f H^o$(298.15 K) = 79.39 kJ/mol – S^o(298.15 K) = 158.75 J/mol-K – p^o = 0.1 MPa



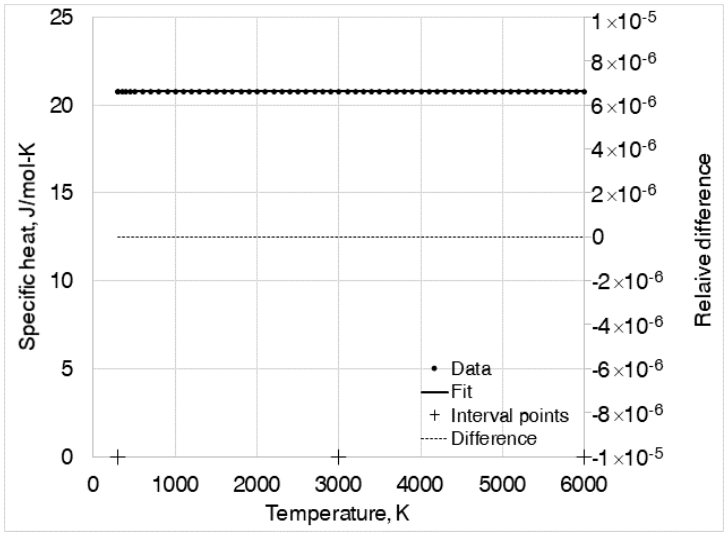


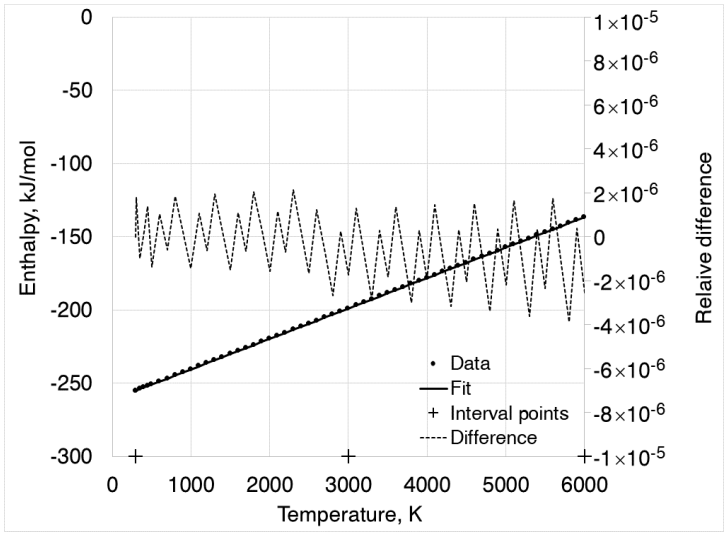


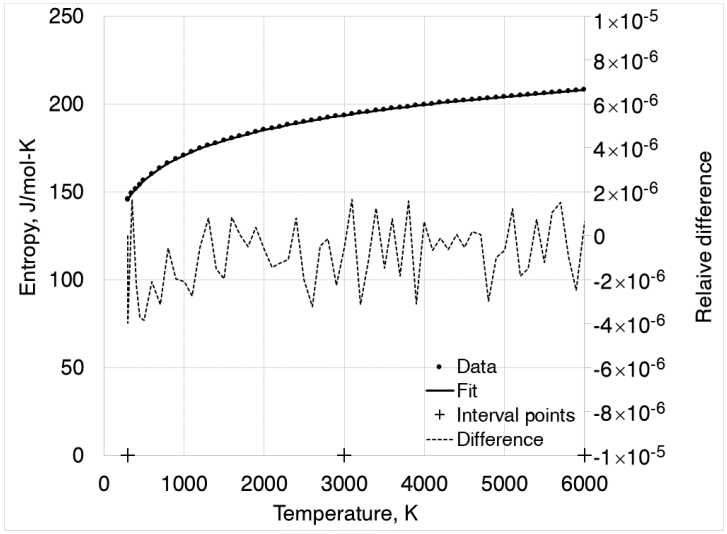


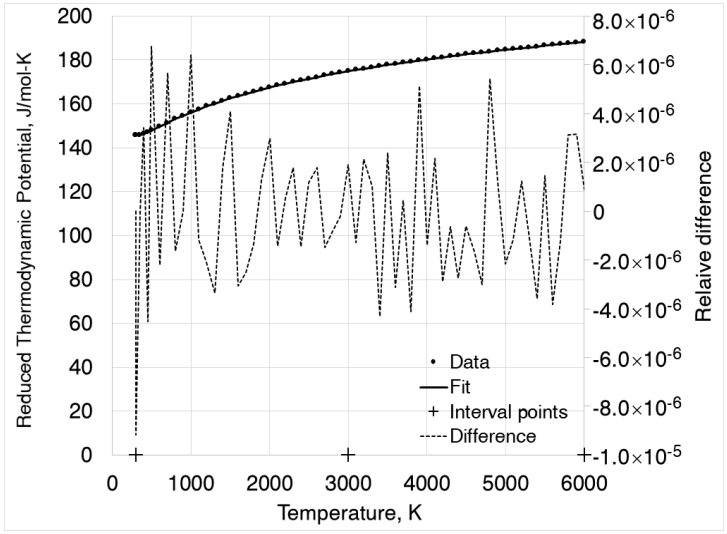
Fluorine, Ion (F^-) – F1-(g)

MW = 18.999 g/mol – $\Delta_f H^o$(298.15 K) = -255.079 kJ/mol – S^o(298.15 K) = 145.576 J/mol-K – p^o = 0.1 MPa



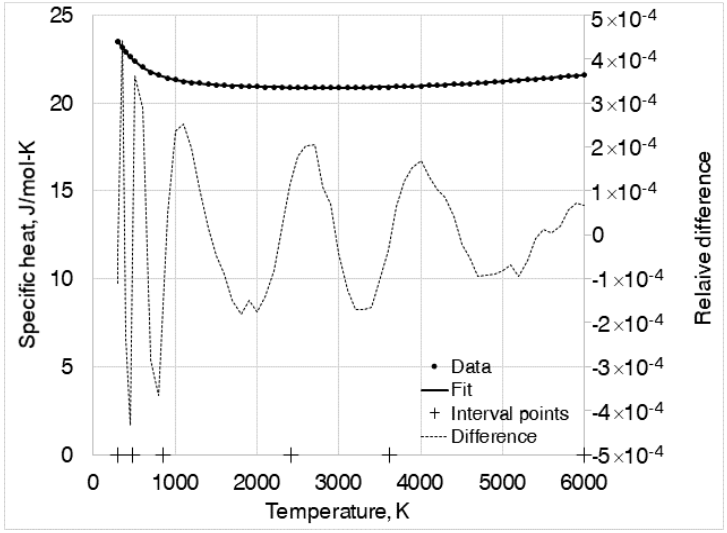


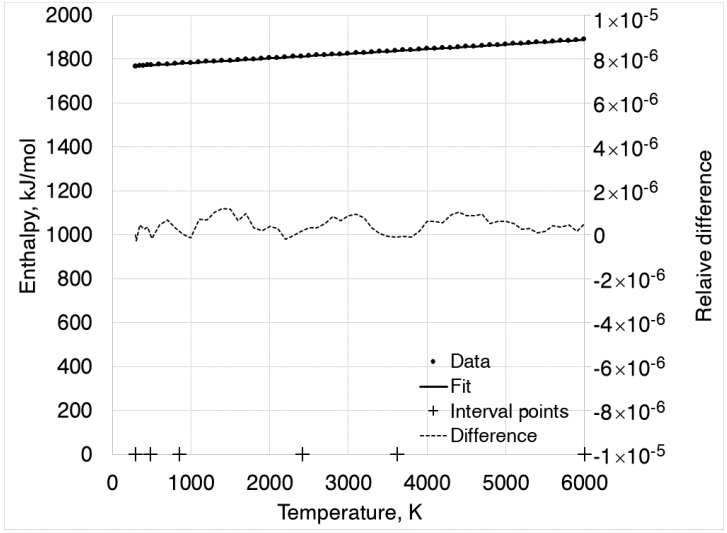


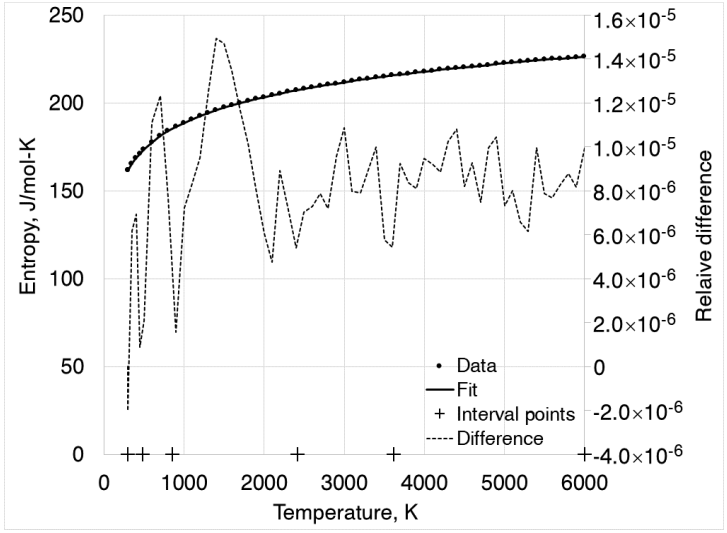


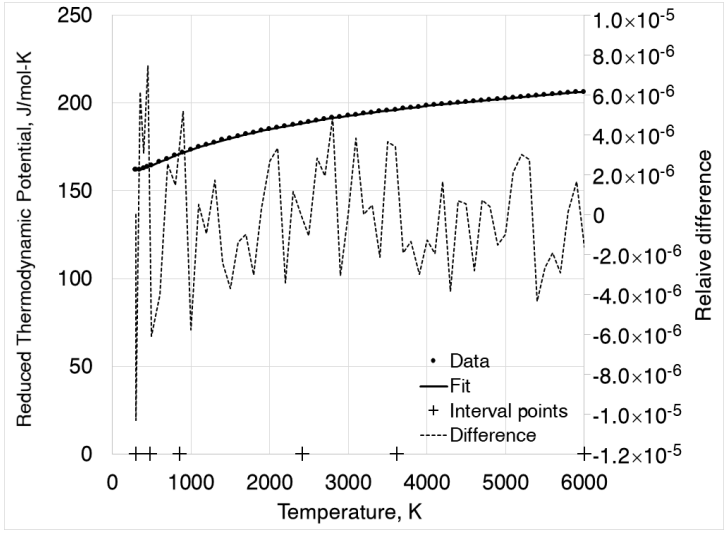
Fluorine, Ion (F^+) – F1+(g)

MW = 18.9979 g/mol – $\Delta_f H^o$(298.15 K) = 1766.826 kJ/mol – S^o(298.15 K) = 161.728 J/mol-K – p^o = 0.1 MPa



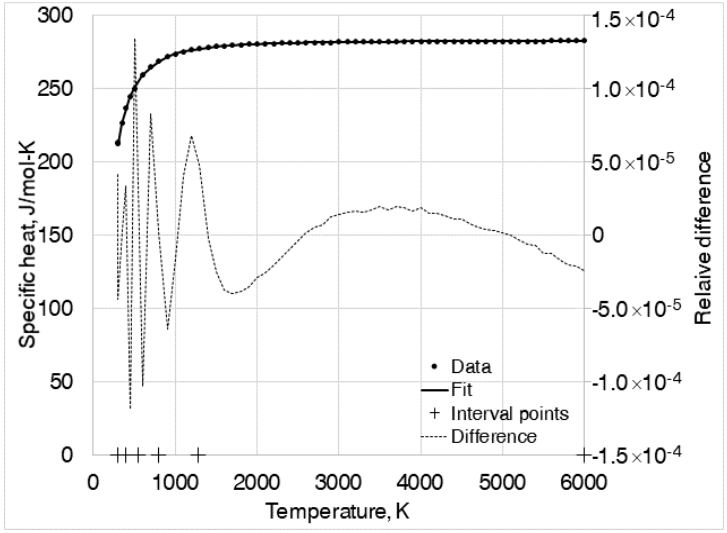


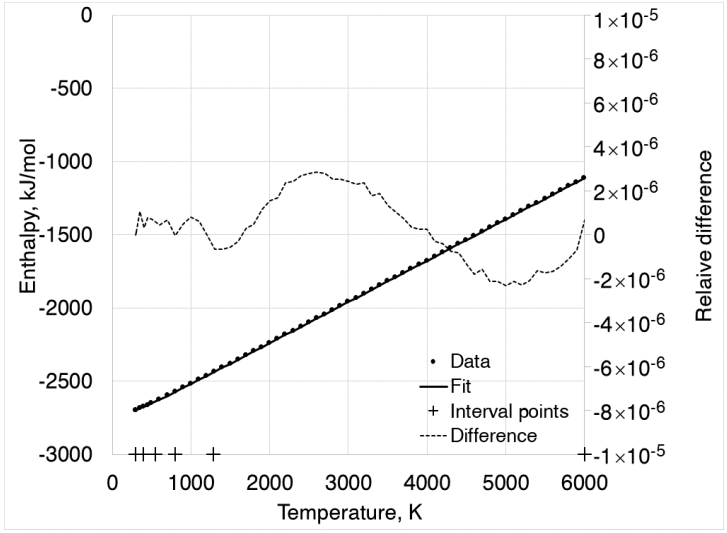


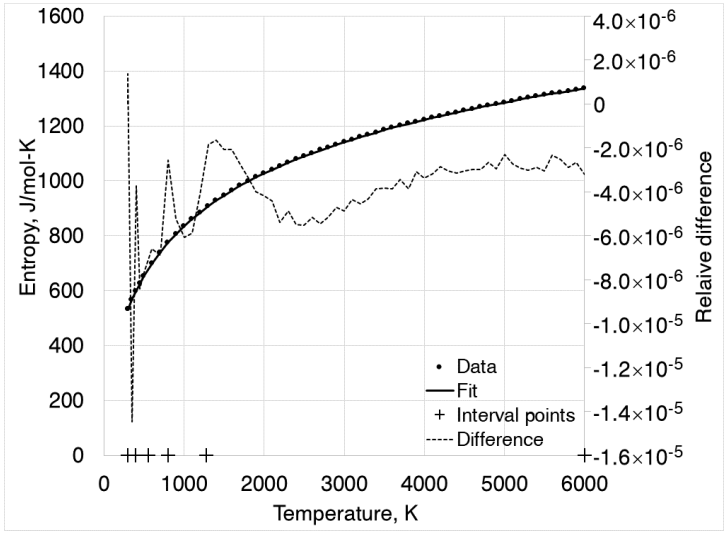


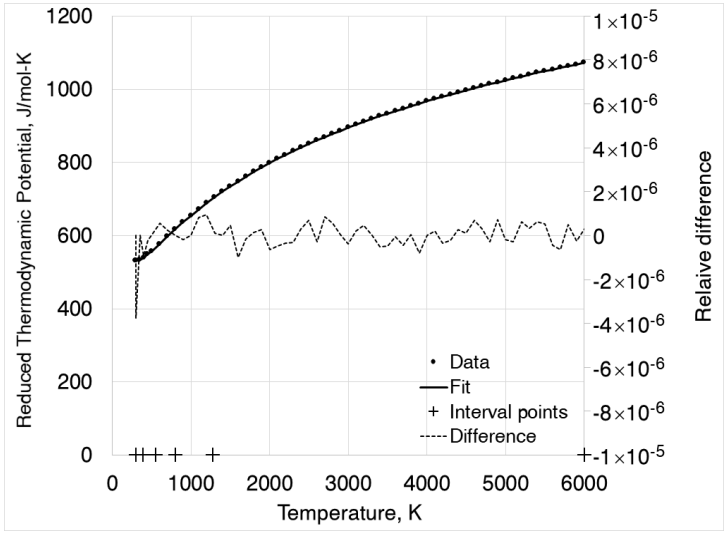
Molybdenum Fluoride ($(MoF_5)_2$) – F10Mo2(g)

MW = 381.86 g/mol – $\Delta_f H^o$(298.15 K) = -2697.843 kJ/mol – S^o(298.15 K) = 531.554 J/mol-K – p^o = 0.1 MPa



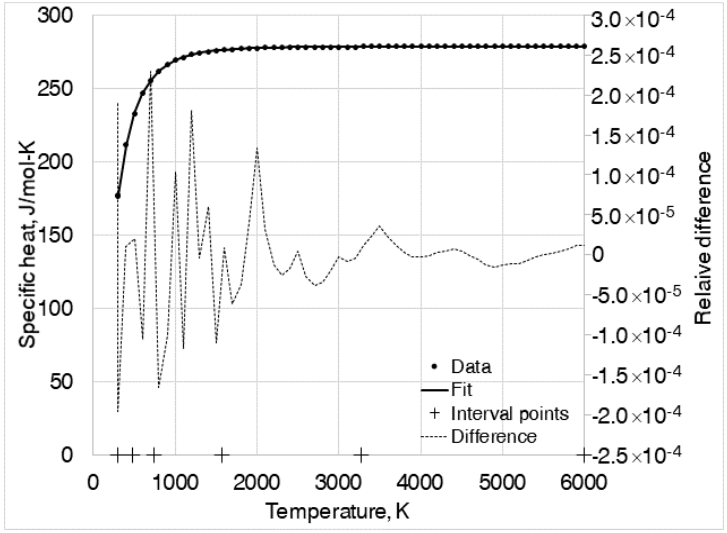


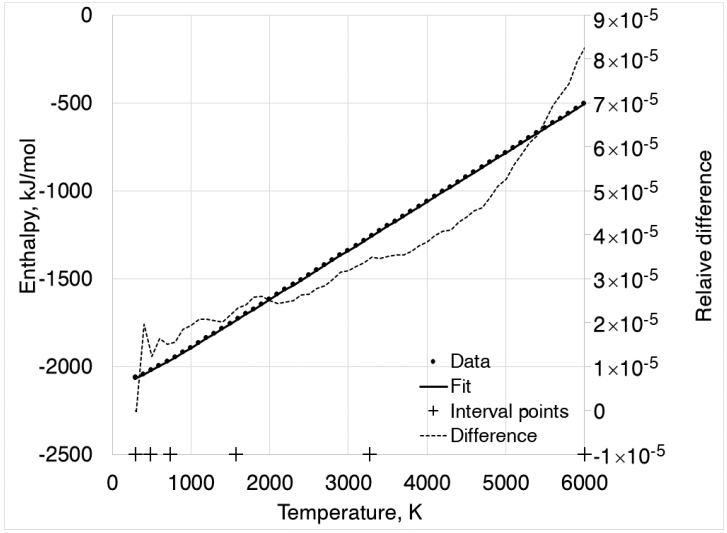


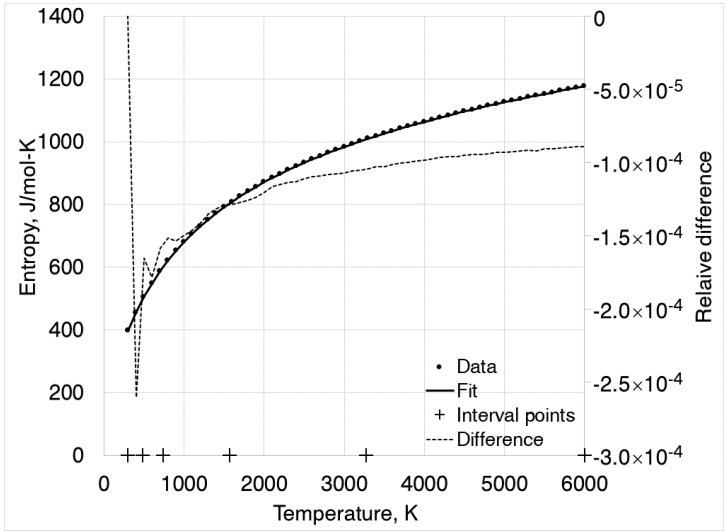


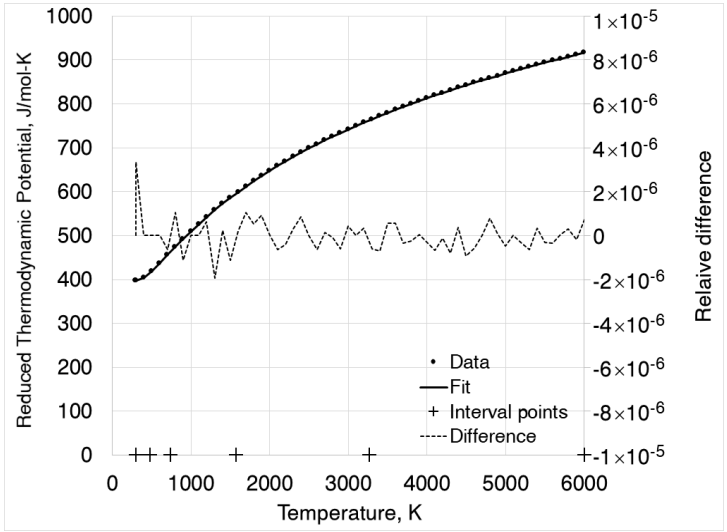
Sulfur Fluoride (S_2F_{10}) – F10S2(g)

MW = 254.1 g/mol – $\Delta_f H^o$(298.15 K) = -2064.386 kJ/mol – S^o(298.15 K) = 396.943 J/mol-K – p^o = 0.1 MPa



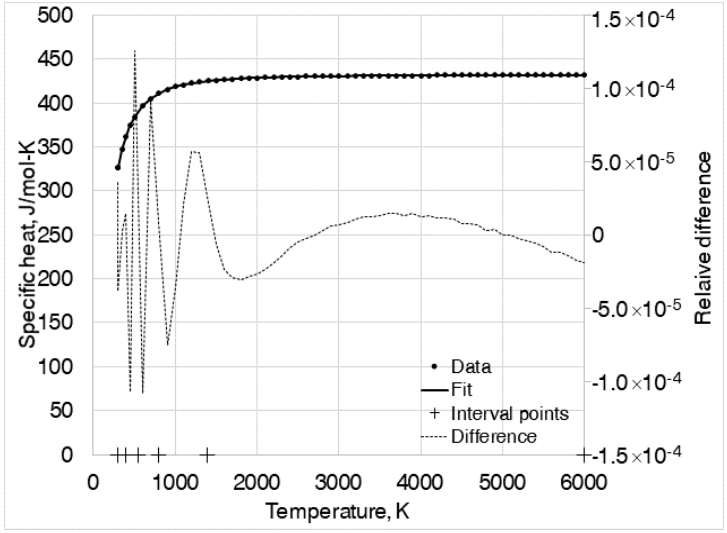


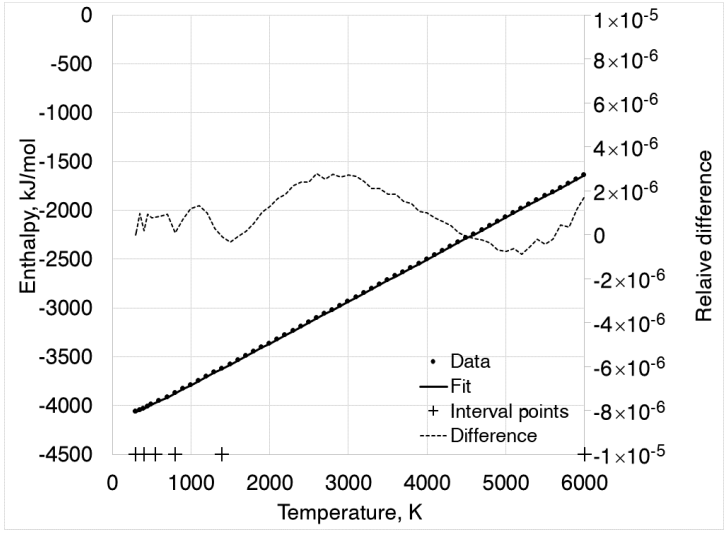


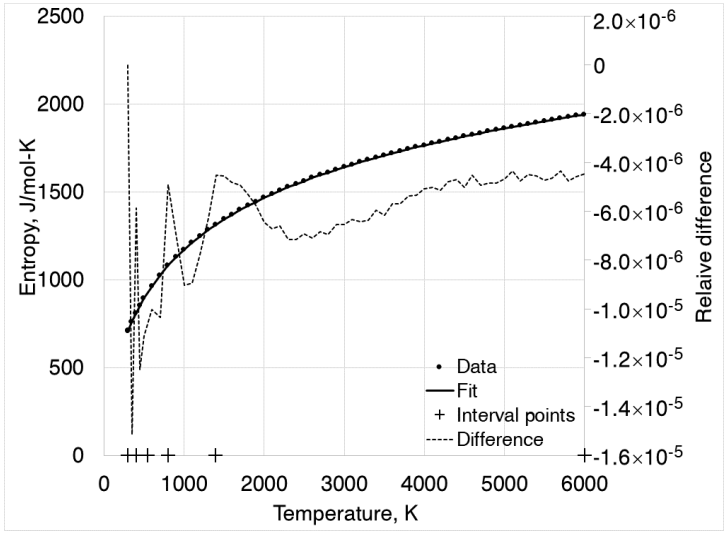


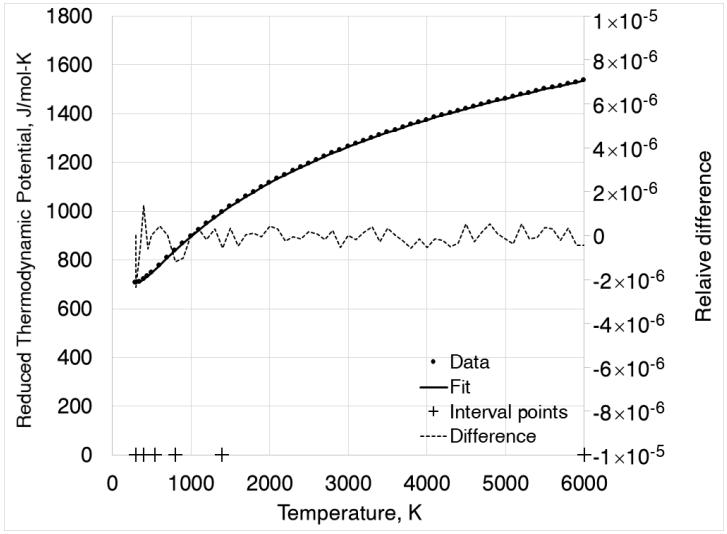
Molybdenum Fluoride ($(MoF_5)_3$) – F15Mo3(g)

MW = 572.8 g/mol – $\Delta_f H^o$(298.15 K) = -4065.593 kJ/mol – S^o(298.15 K) = 706.844 J/mol-K – p^o = 0.1 MPa



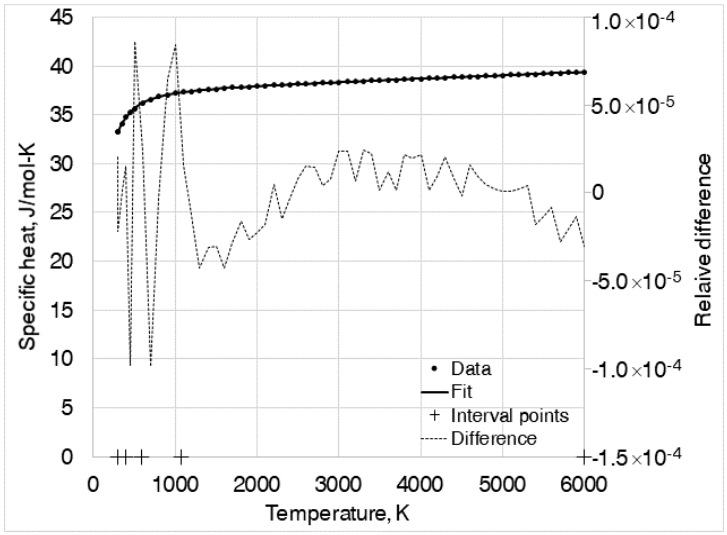


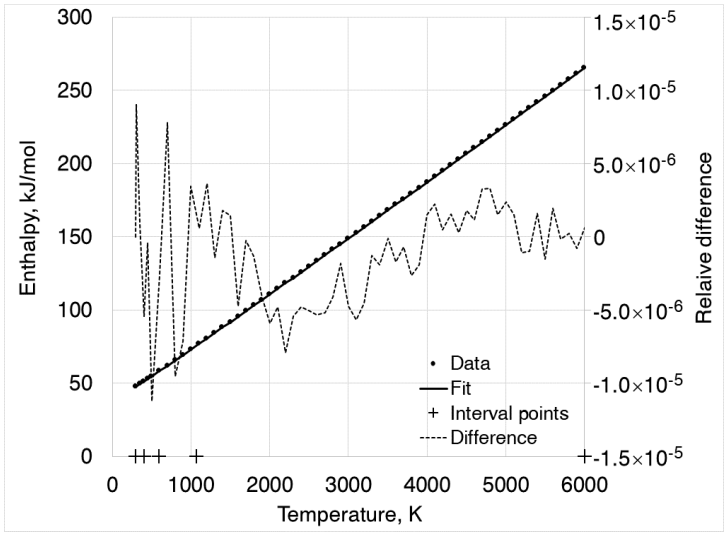


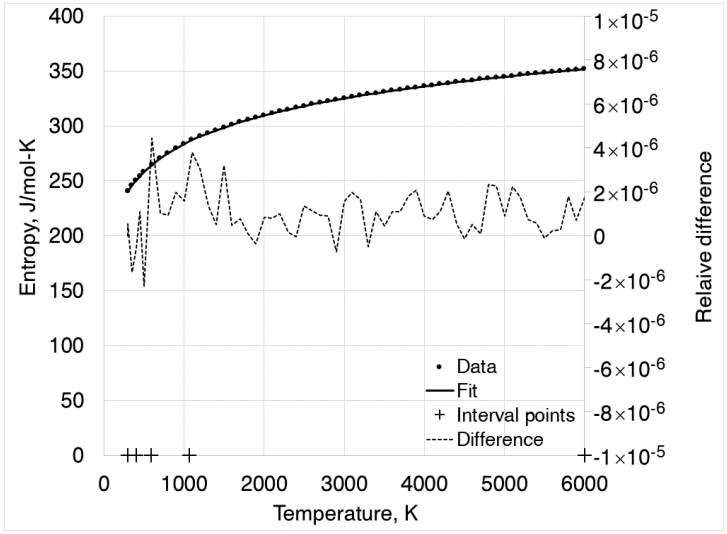


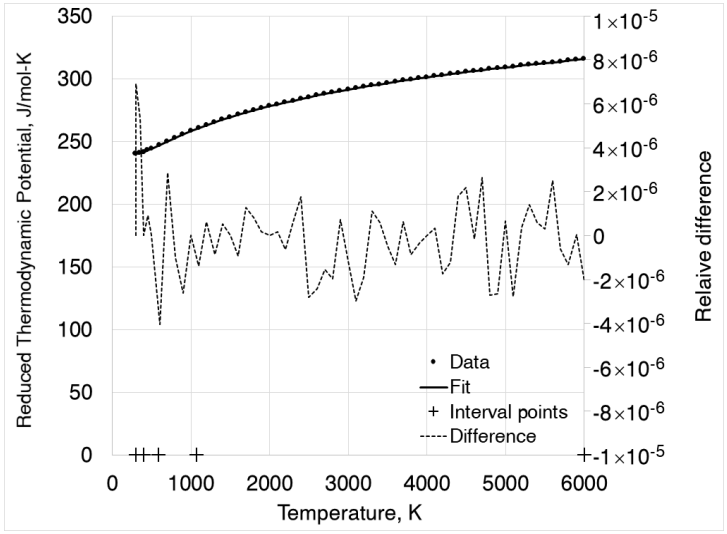
Iron Fluoride (FeF) – F1Fe1(g)

MW = 74.845 g/mol – $\Delta_f H^o$(298.15 K) = 47.698 kJ/mol – S^o(298.15 K) = 240.337 J/mol-K – p^o = 0.1 MPa



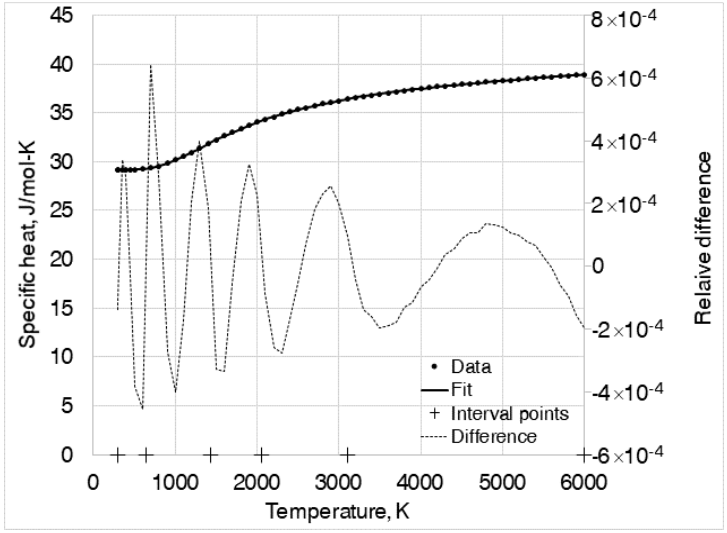


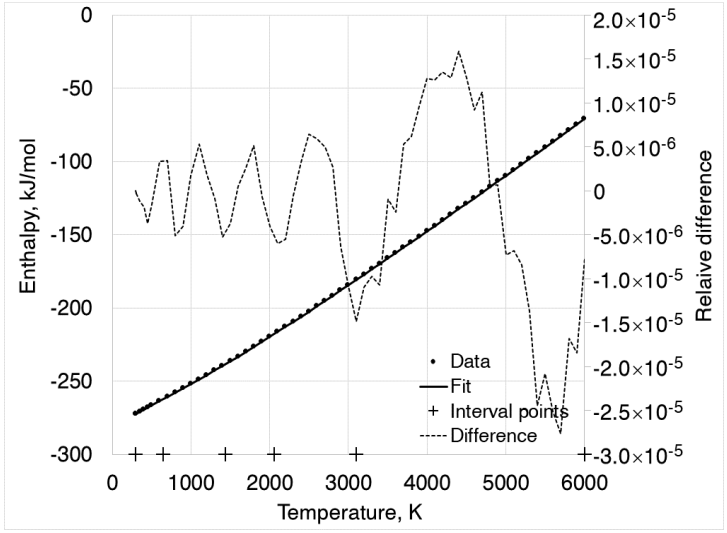


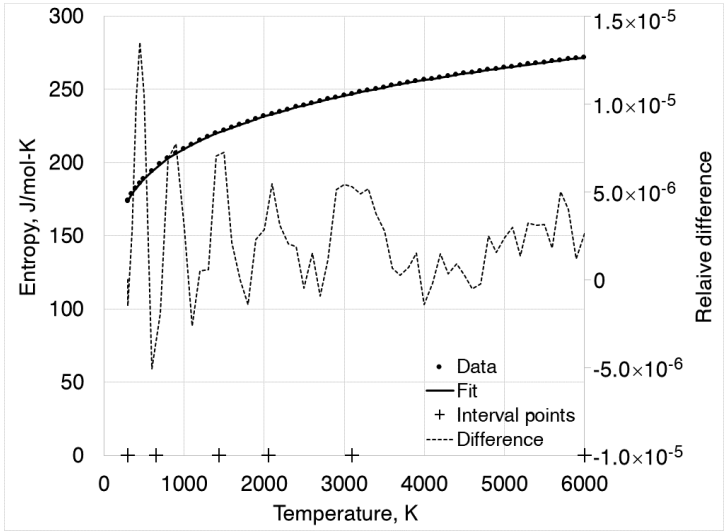


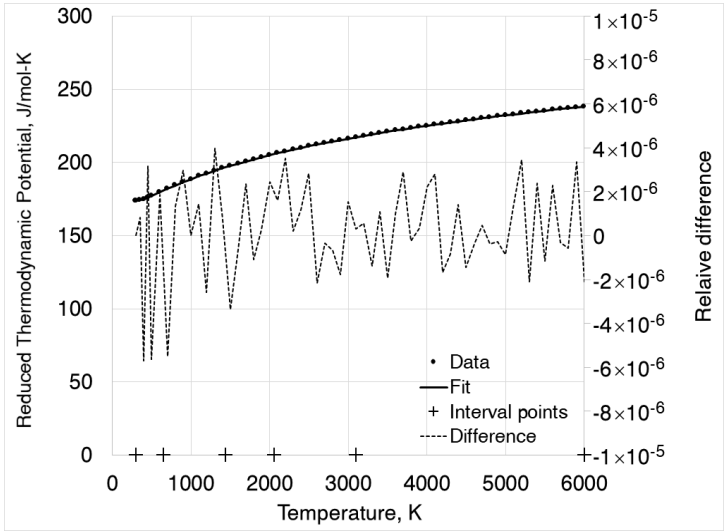
Hydrogen Fluoride (HF) – F1H1(g)

MW = 20.0063 g/mol – $\Delta_f H^o$(298.15 K) = -272.546 kJ/mol – S^o(298.15 K) = 173.78 J/mol-K – p^o = 0.1 MPa



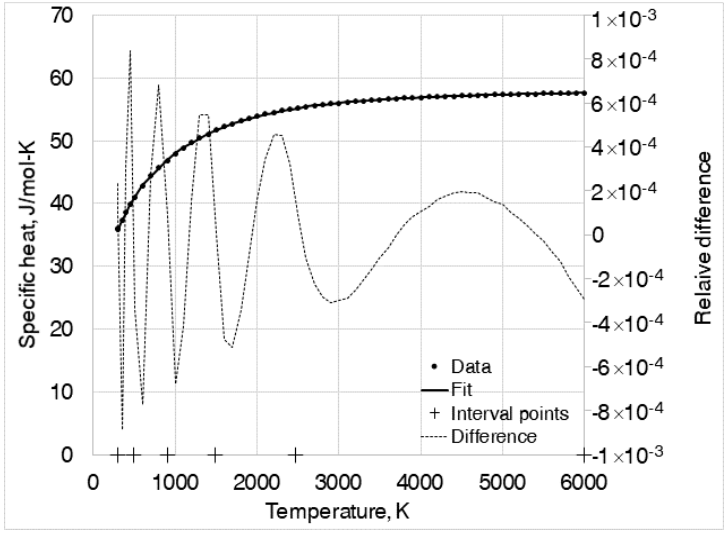


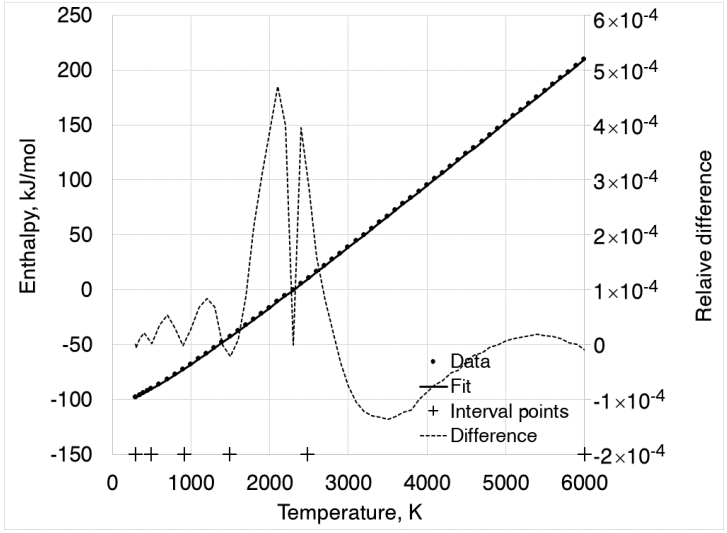


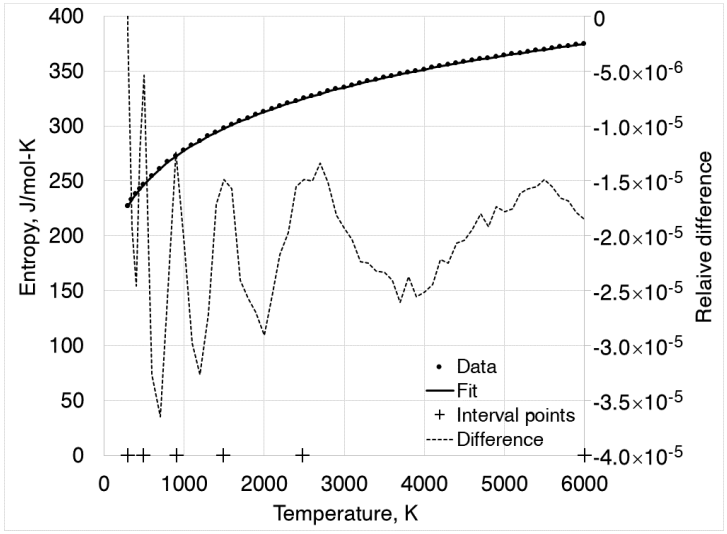


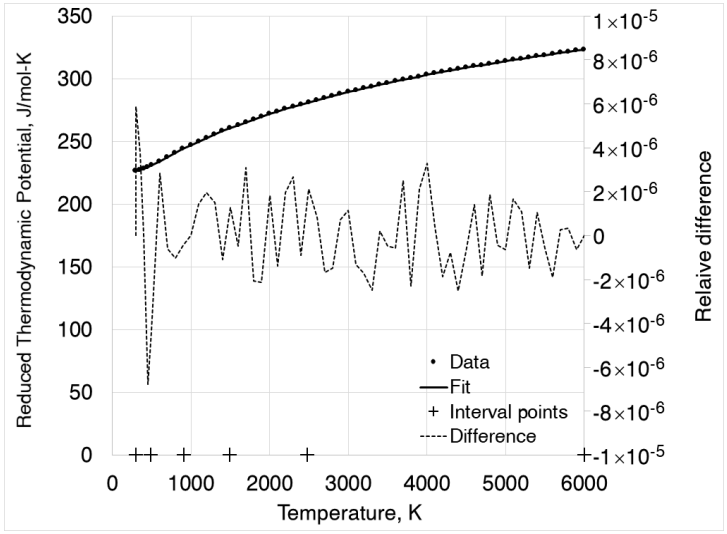
Hypofluorous Acid (HOF) – F1H1O1(g)

MW = 36.0057 g/mol – $\Delta_f H^o$(298.15 K) = -98.324 kJ/mol – S^o(298.15 K) = 226.772 J/mol-K – p^o = 0.1 MPa



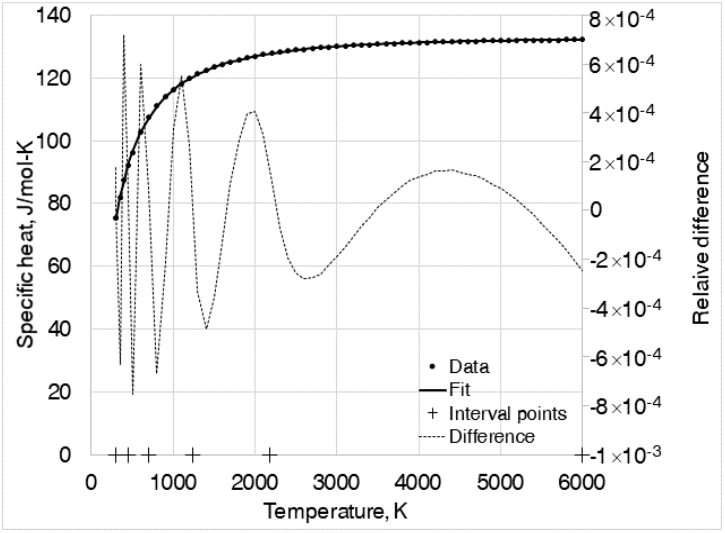


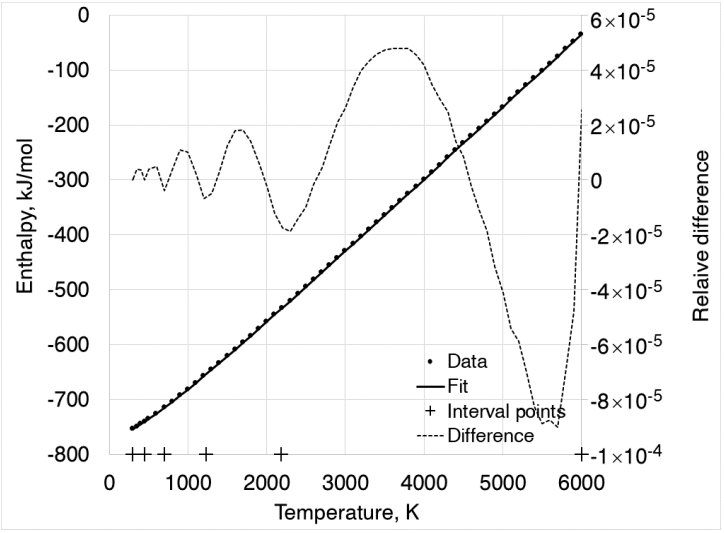


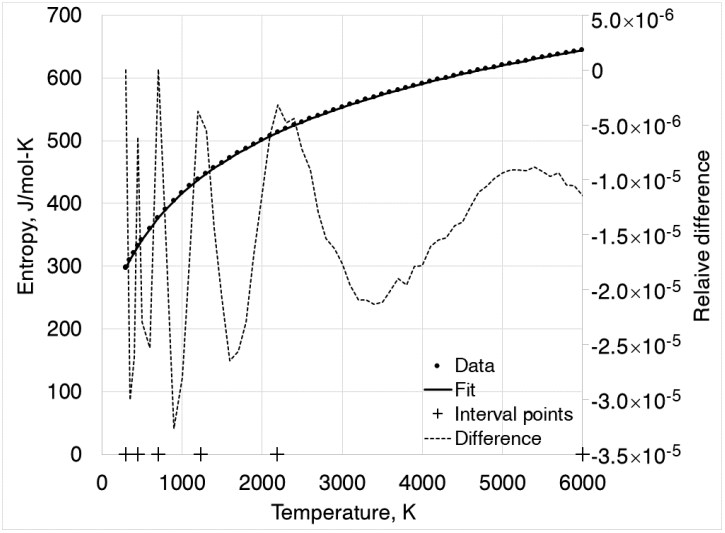


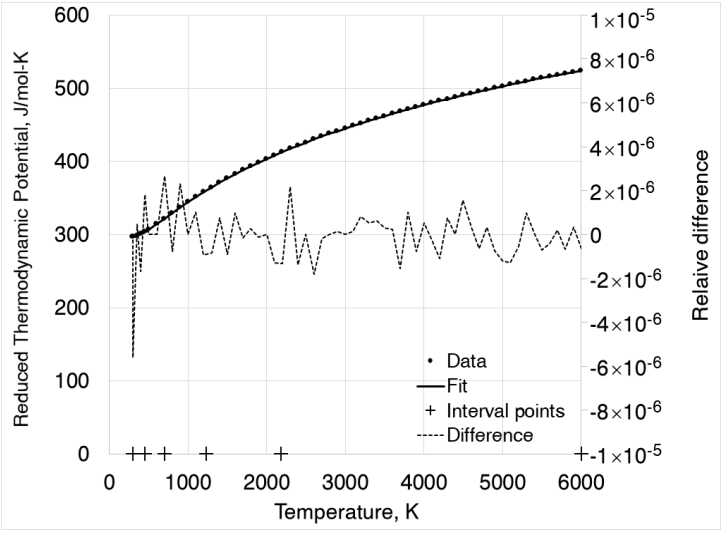
Fluorosulfuric Acid (HSO_3F) – F1H1O3S1(g)

MW = 100.06 g/mol – $\Delta_f H^o$(298.15 K) = -753.12 kJ/mol – S^o(298.15 K) = 297.267 J/mol-K – p^o = 0.1 MPa



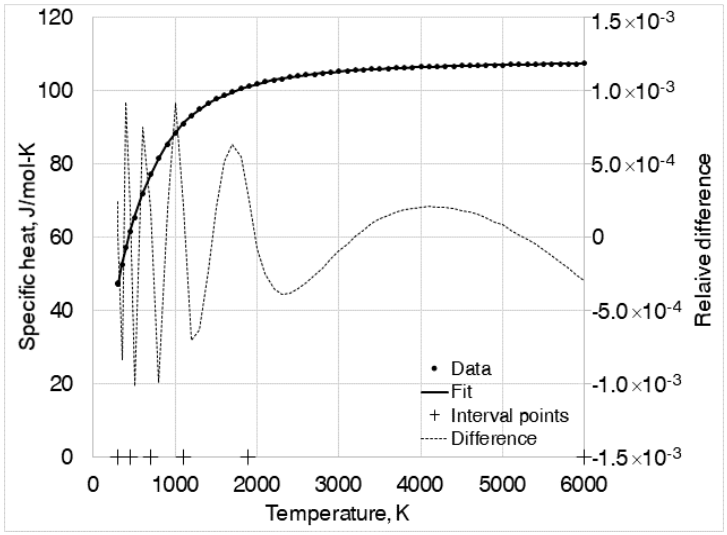


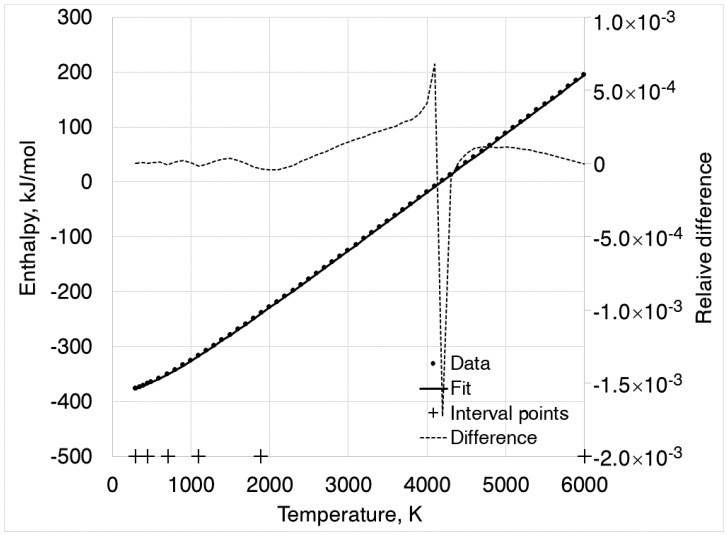


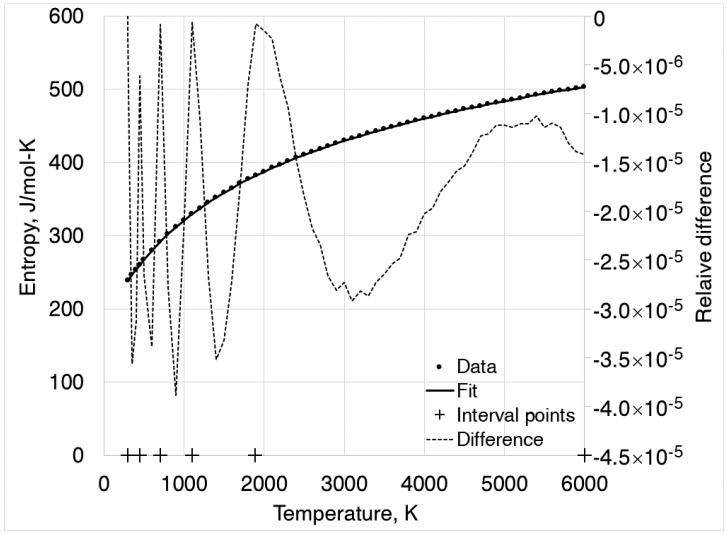


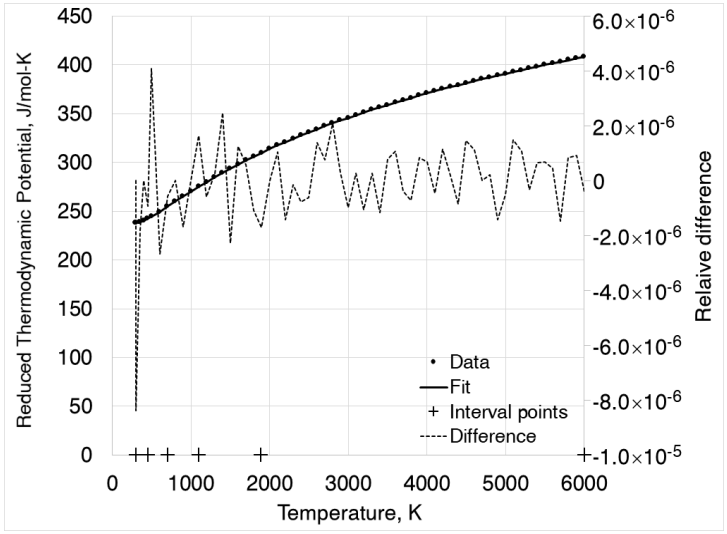
Fluorosilane (SiH_3F) – F1H3Si1(g)

MW = 50.1077 g/mol – $\Delta_f H^o$(298.15 K) = -376.56 kJ/mol – S^o(298.15 K) = 238.402 J/mol-K – p^o = 0.1 MPa



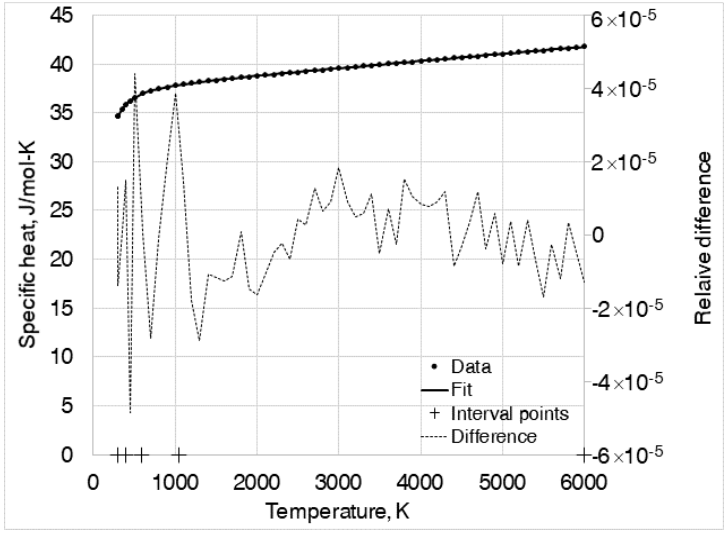


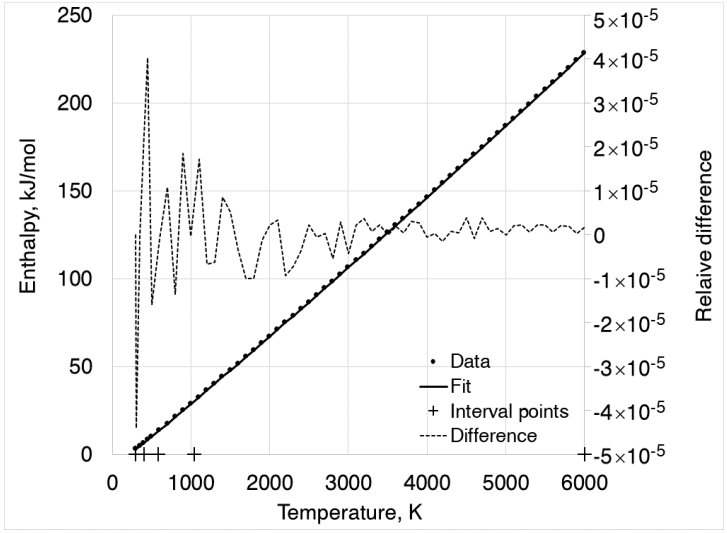


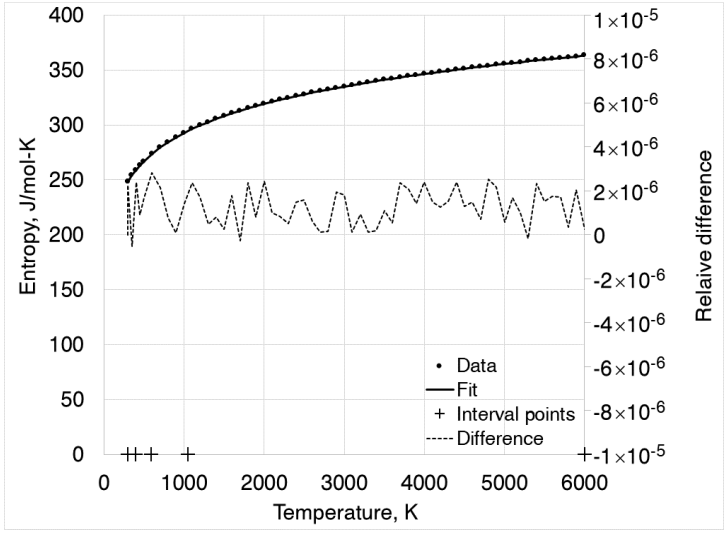


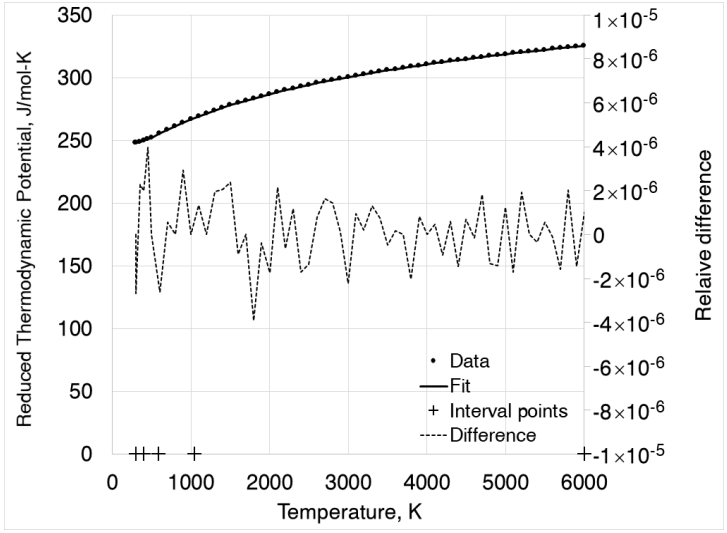
Mercury Fluoride (HgF) – F1Hg1(g)

MW = 219.59 g/mol – $\Delta_f H^o$(298.15 K) = 2.929 kJ/mol – S^o(298.15 K) = 248.373 J/mol-K – p^o = 0.1 MPa



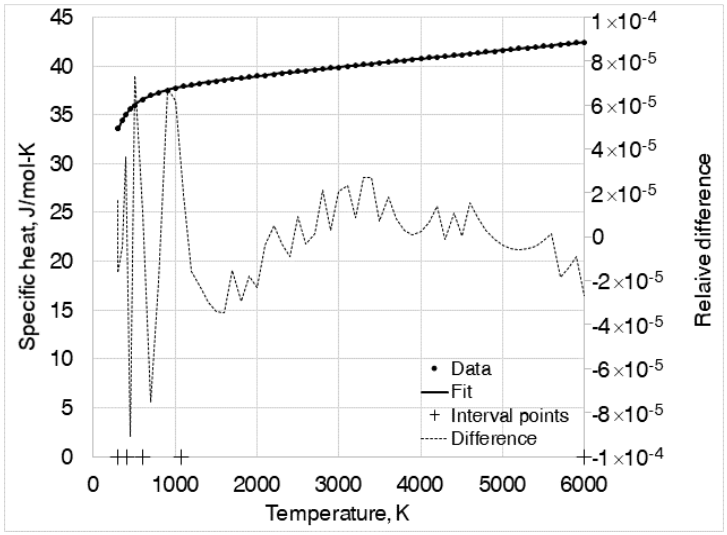


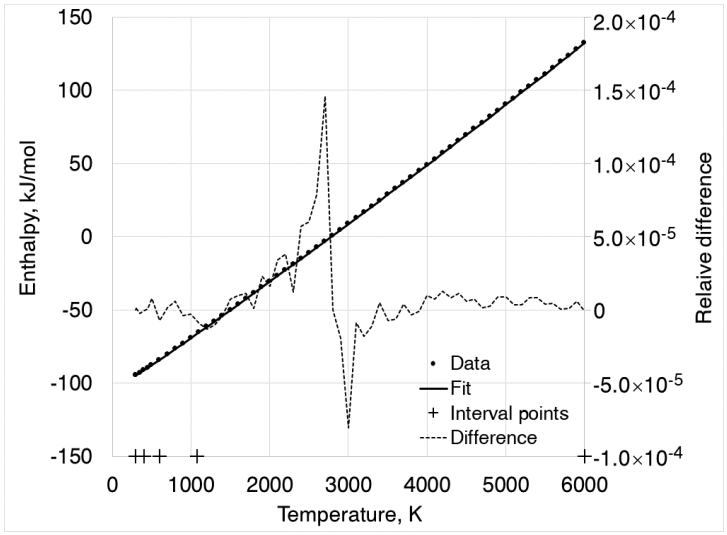


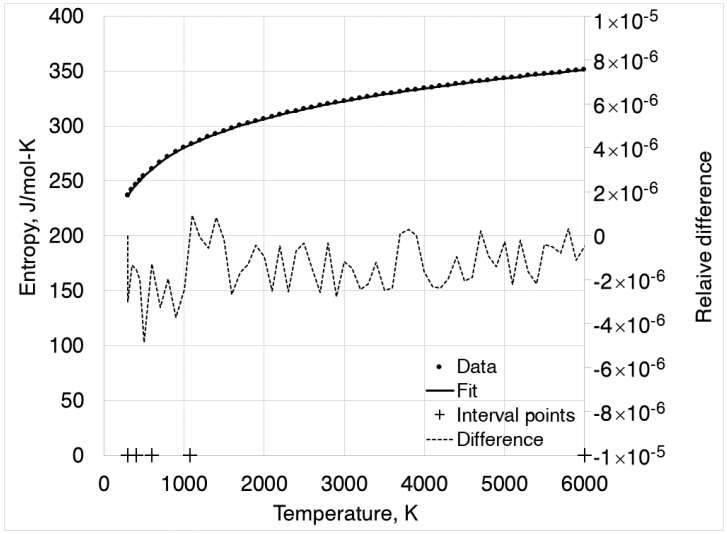


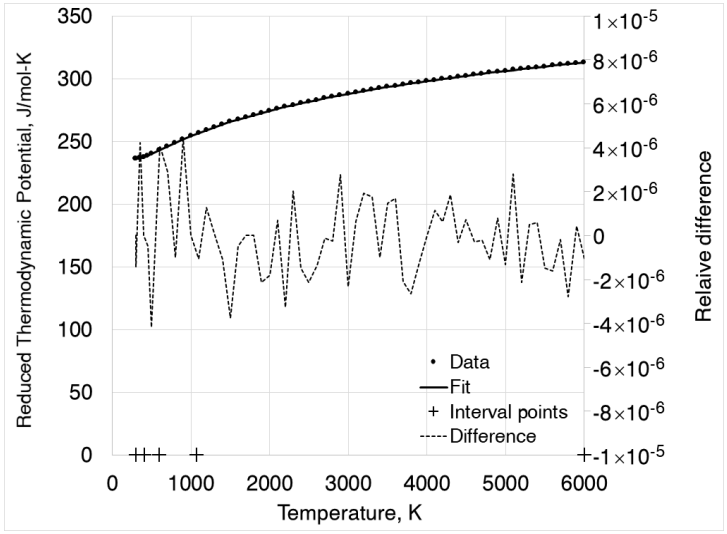
Iodine Fluoride (IF) – F1I1(g)

MW = 145.9029 g/mol – $\Delta_f H^o$(298.15 K) = -94.759 kJ/mol – S^o(298.15 K) = 236.296 J/mol-K – p^o = 0.1 MPa



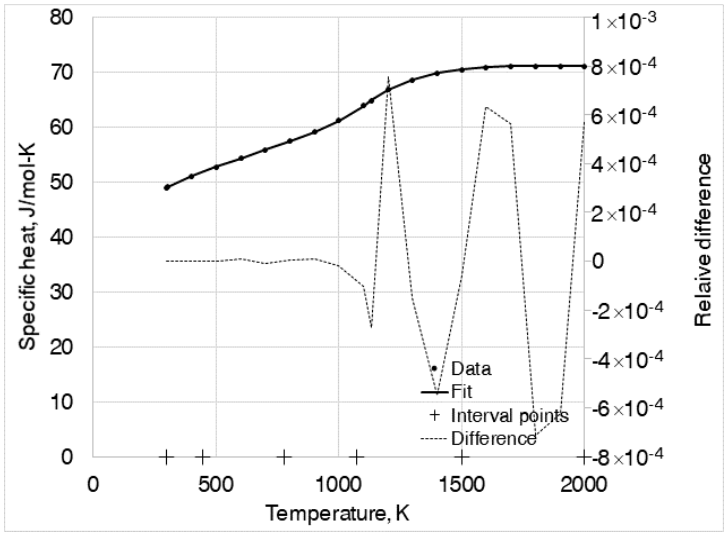


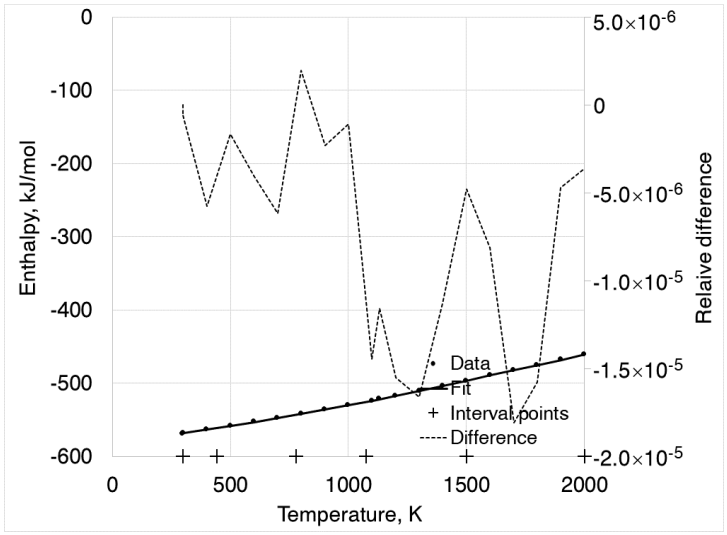


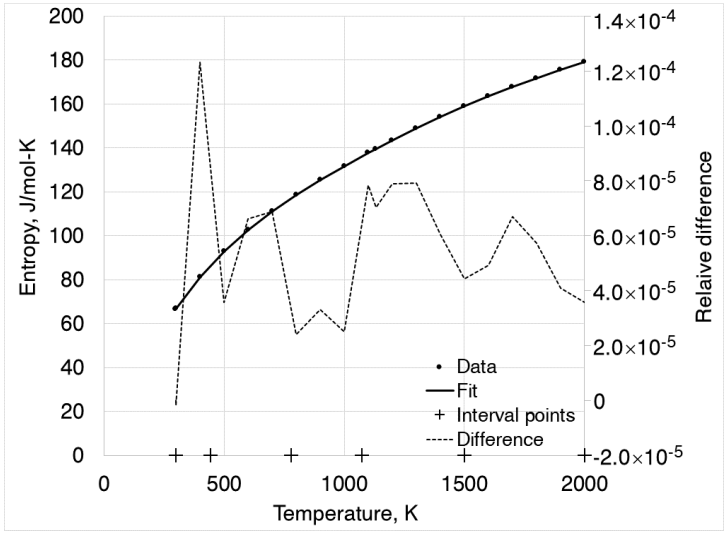


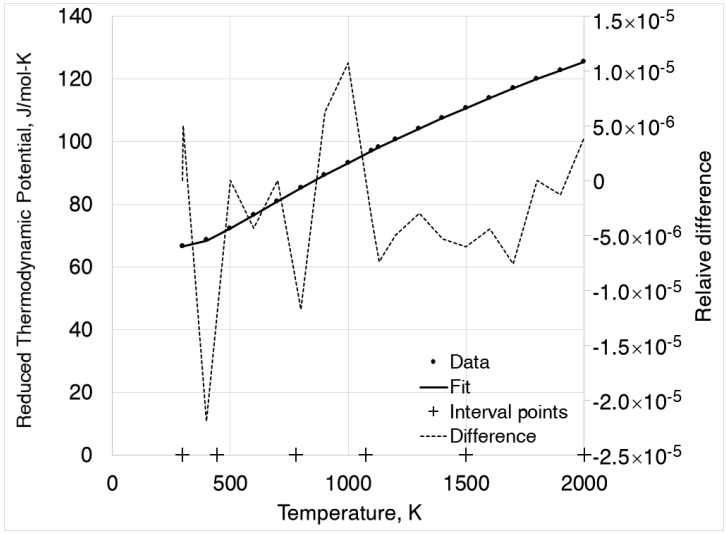
Potassium Fluoride (KF) – F1K1(cr)

MW = 58.0967 g/mol – $\Delta_f H^o$(298.15 K) = -568.606 kJ/mol – S^o(298.15 K) = 66.547 J/mol-K – p^o = 0.1 MPa



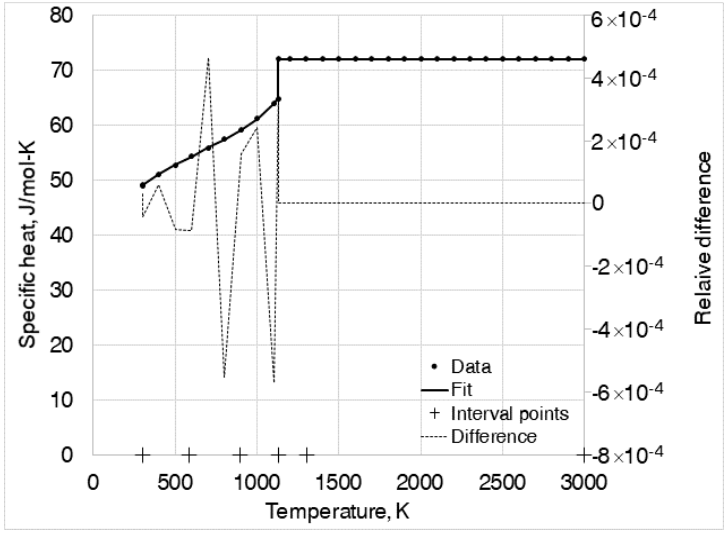


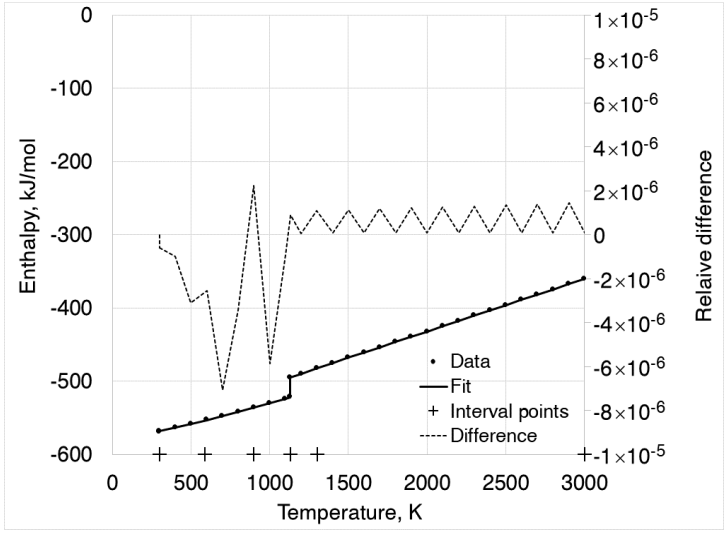


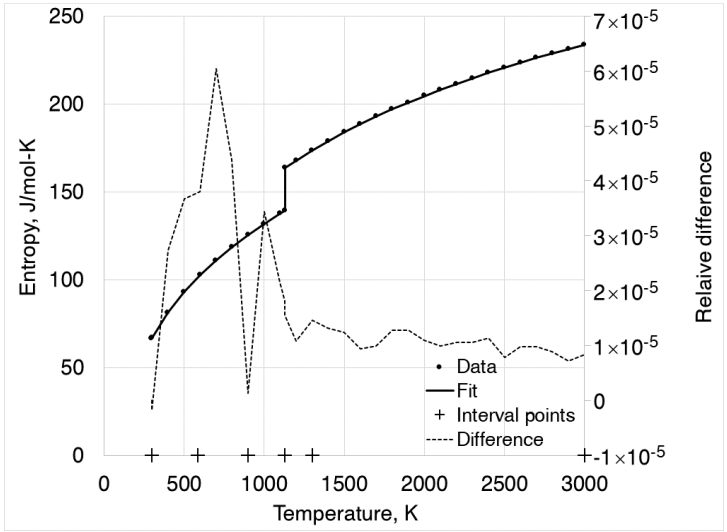


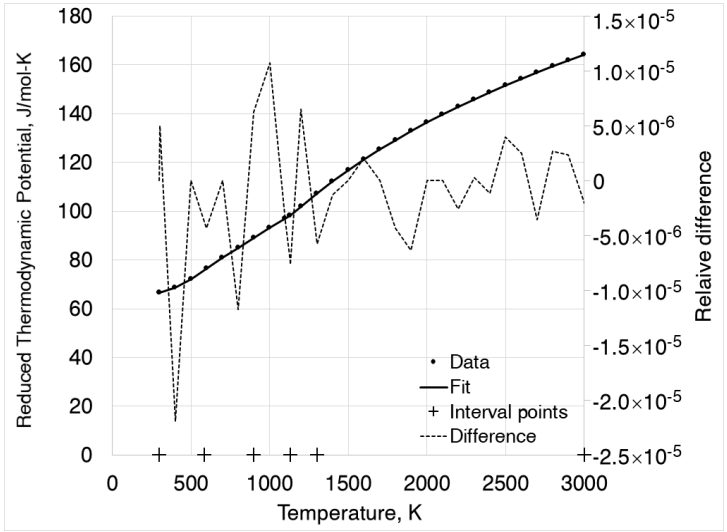
Potassium Fluoride (KF) – F1K1(cr,l)

MW = 58.0967 g/mol – $\Delta_f H^o$(298.15 K) = -568.606 kJ/mol – S^o(298.15 K) = 66.547 J/mol-K – p^o = 0.1 MPa



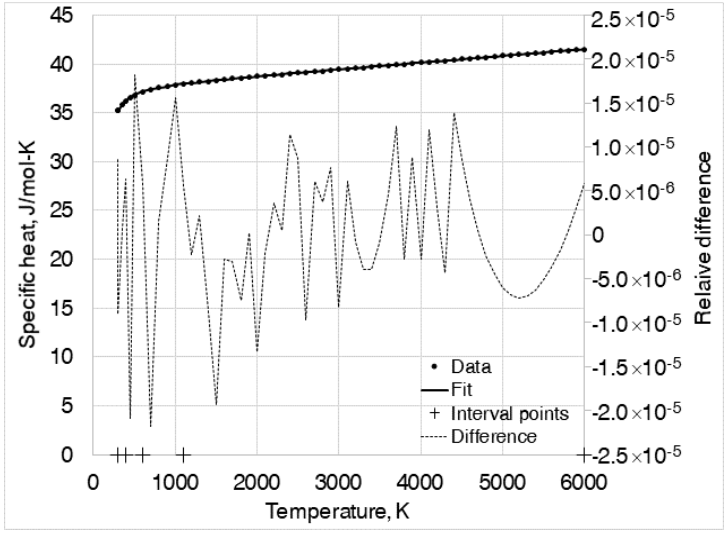


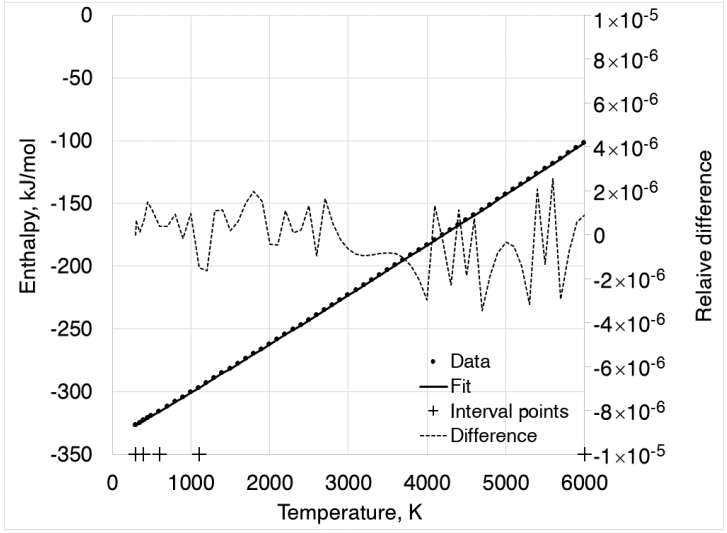


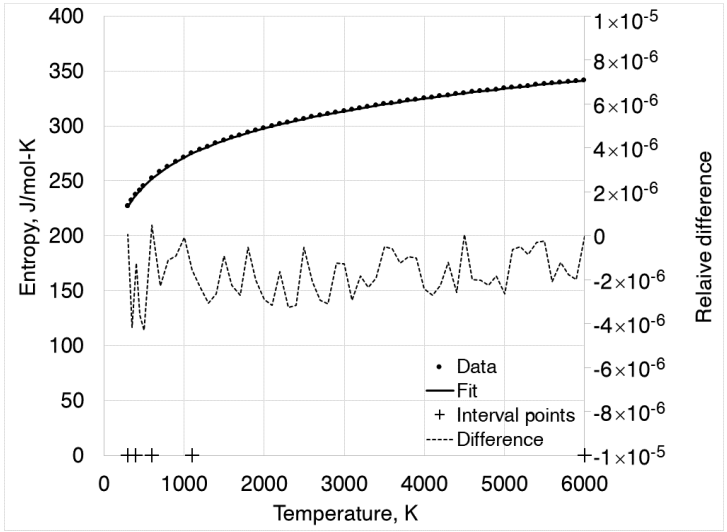


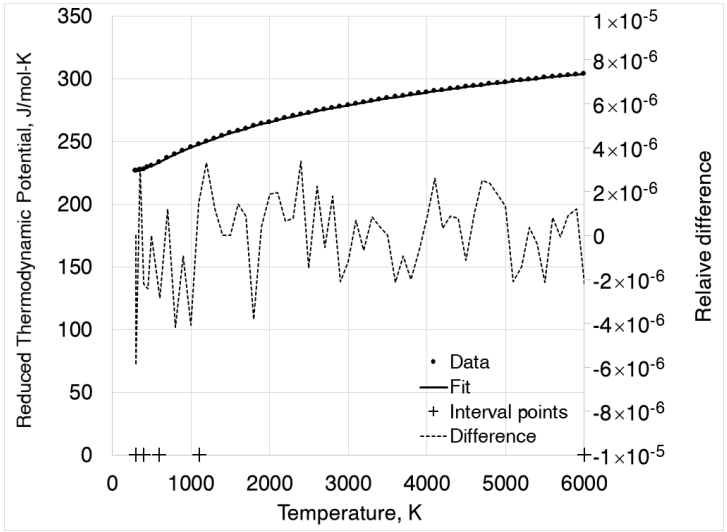
Potassium Fluoride (KF) – F1K1(g)

MW = 58.0967 g/mol – $\Delta_f H^o$(298.15 K) = -326.77 kJ/mol – S^o(298.15 K) = 226.612 J/mol-K – p^o = 0.1 MPa



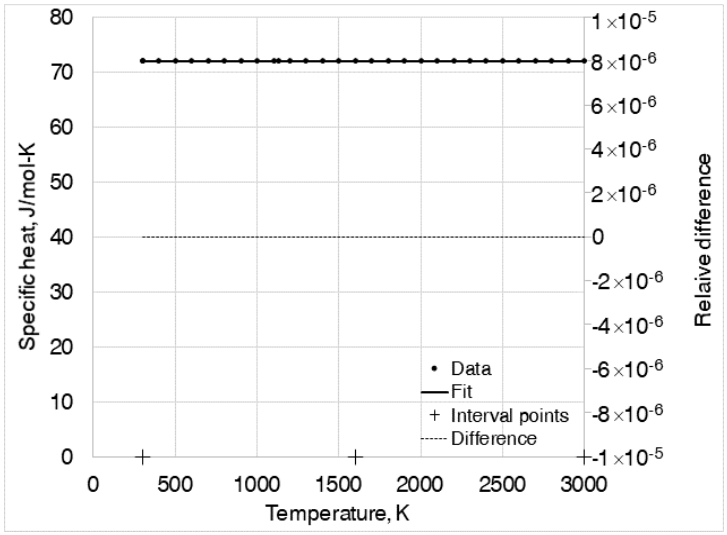


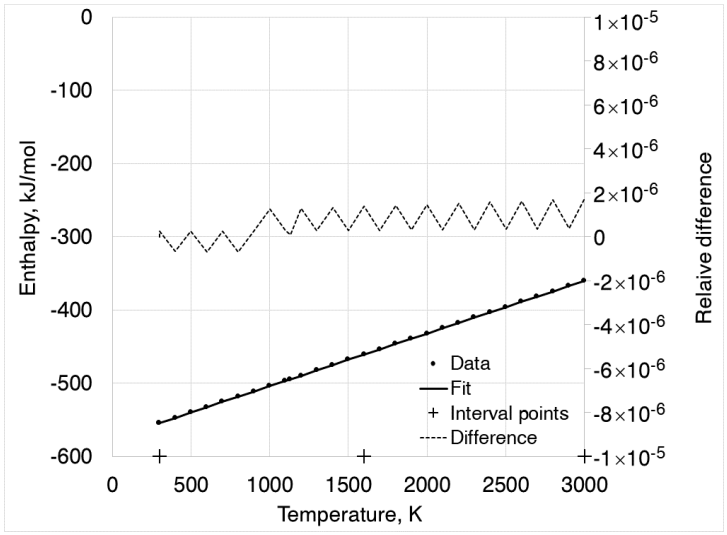


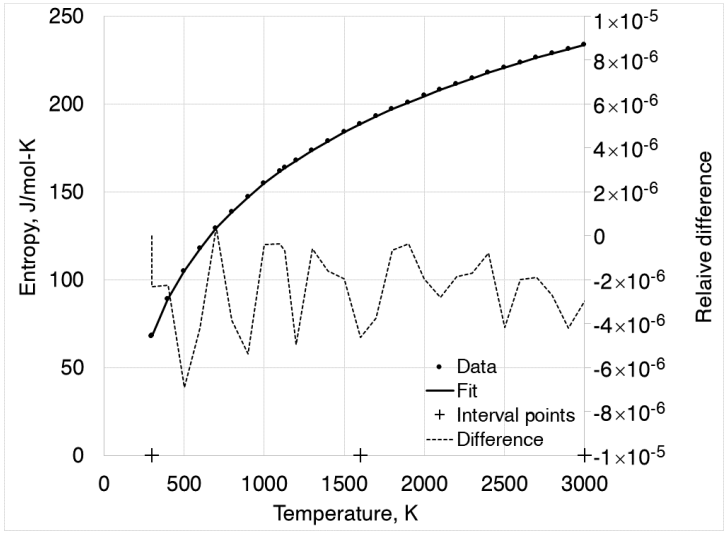


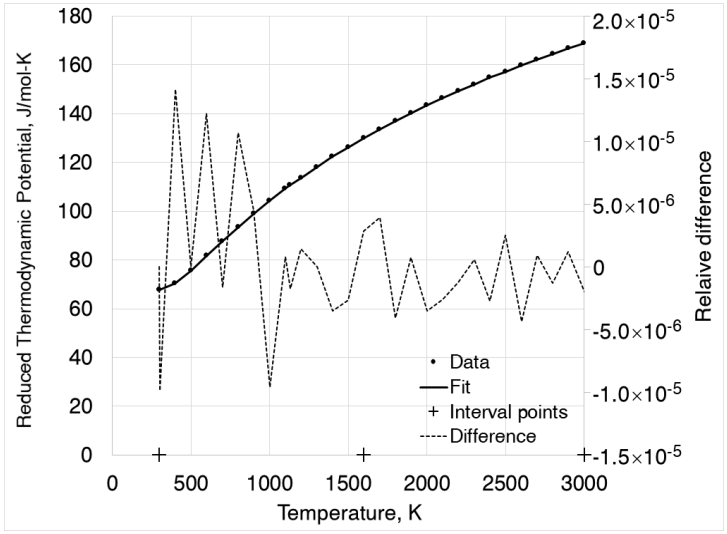
Potassium Fluoride (KF) – F1K1(l)

MW = 58.0967 g/mol – $\Delta_f H^o$(298.15 K) = -554.467 kJ/mol – S^o(298.15 K) = 67.634 J/mol-K – p^o = 0.1 MPa



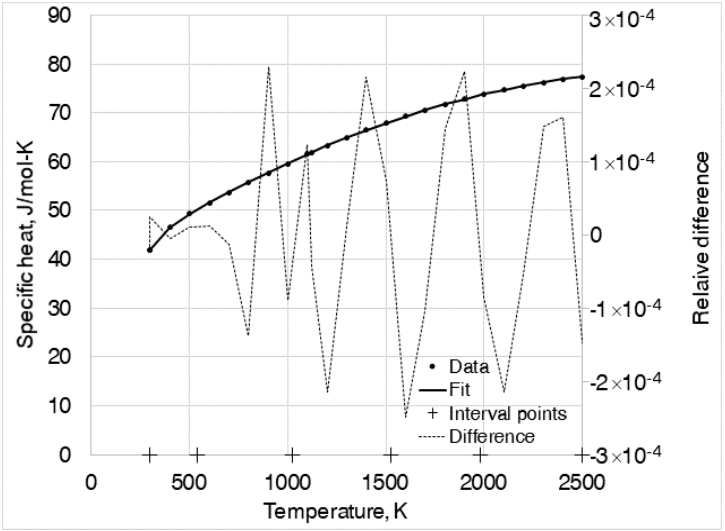


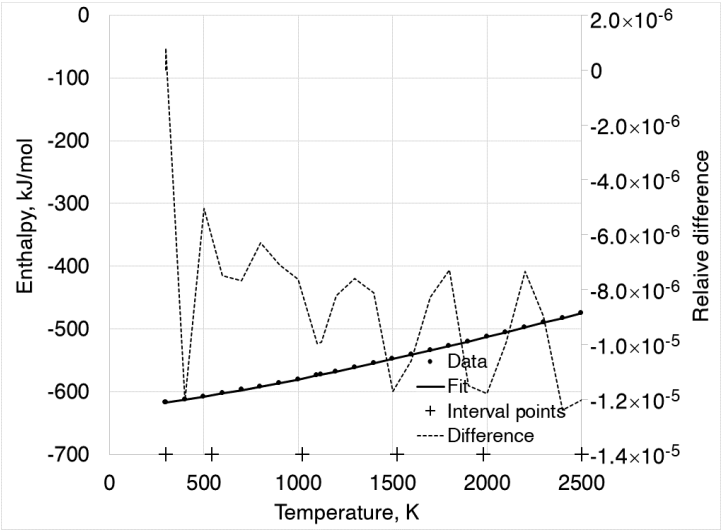


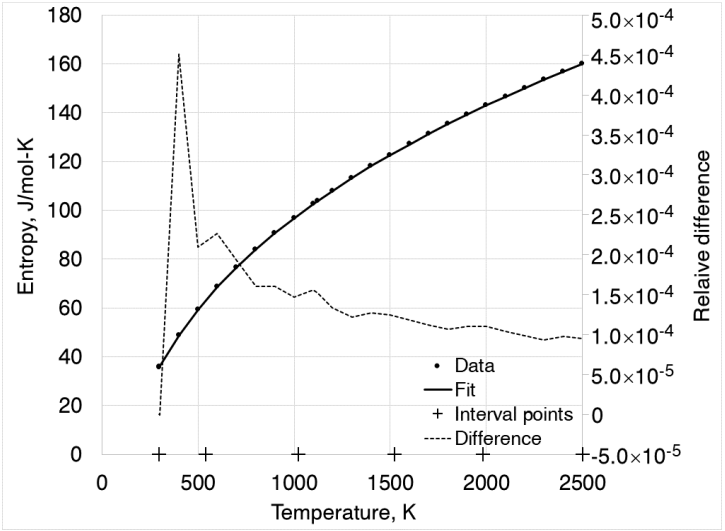


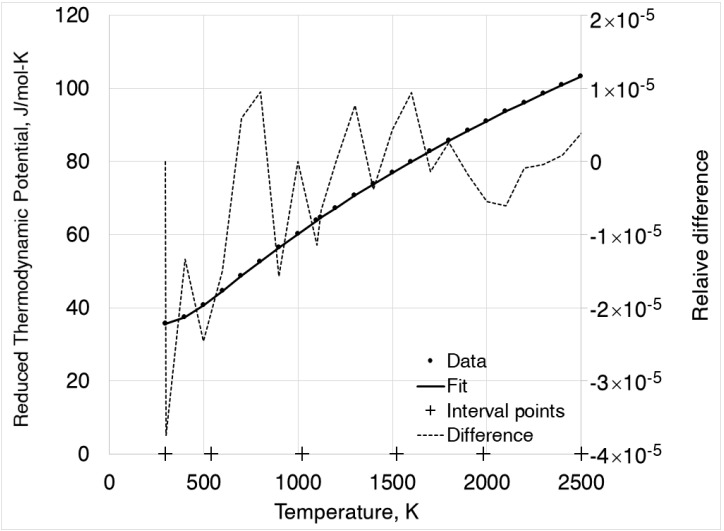
Lithium Fluoride (LiF) – F1Li1(cr)

MW = 25.939 g/mol – $\Delta_f H^o$(298.15 K) = -616.931 kJ/mol – S^o(298.15 K) = 35.66 J/mol-K – p^o = 0.1 MPa



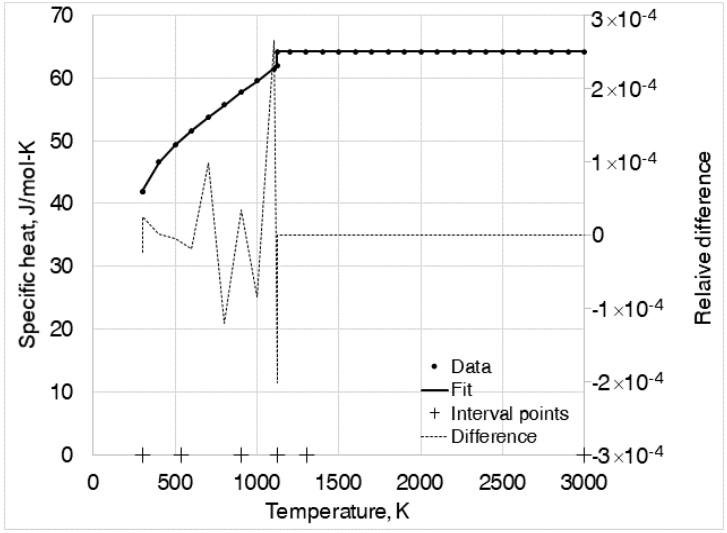


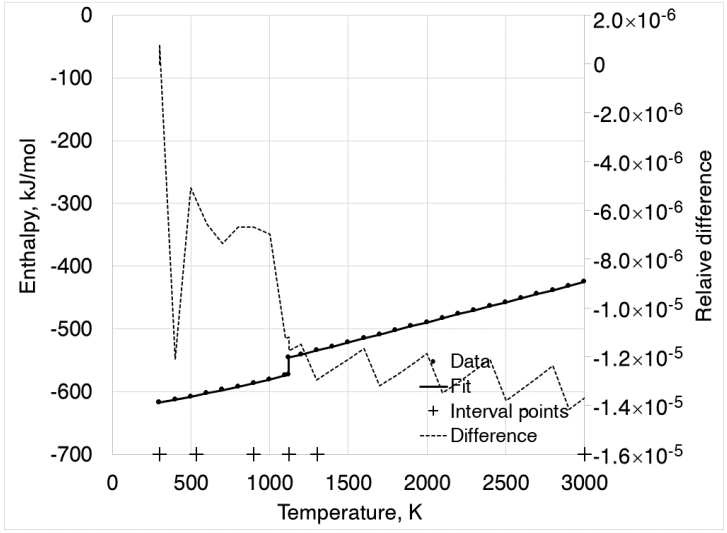


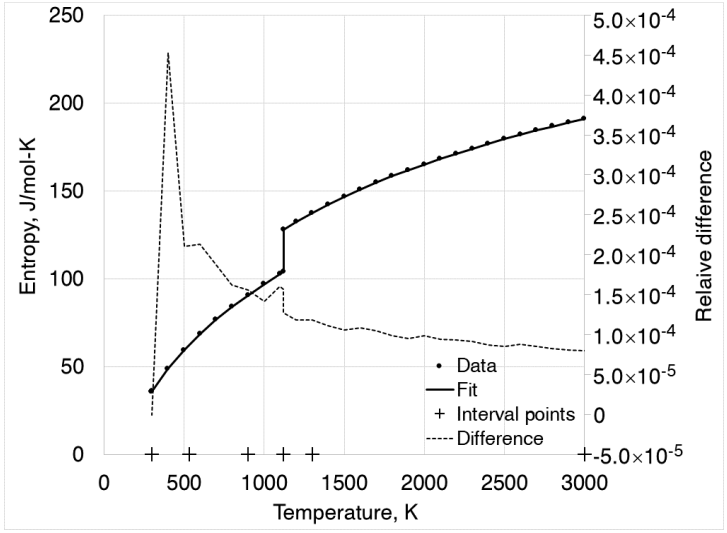


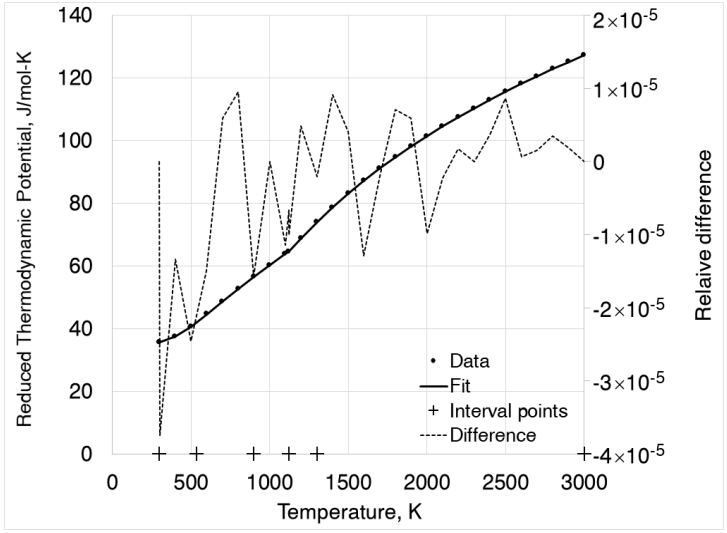
Lithium Fluoride (LiF) – F1Li1(cr,l)

MW = 25.939 g/mol – $\Delta_f H^o$(298.15 K) = -616.931 kJ/mol – S^o(298.15 K) = 35.66 J/mol-K – p^o = 0.1 MPa



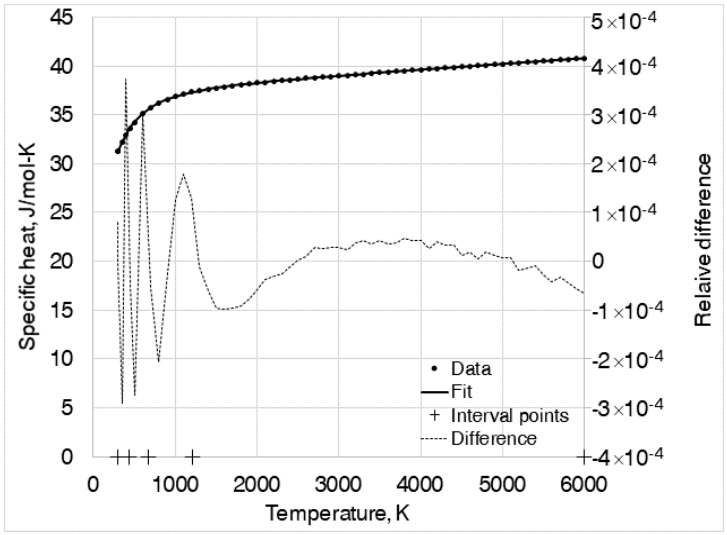


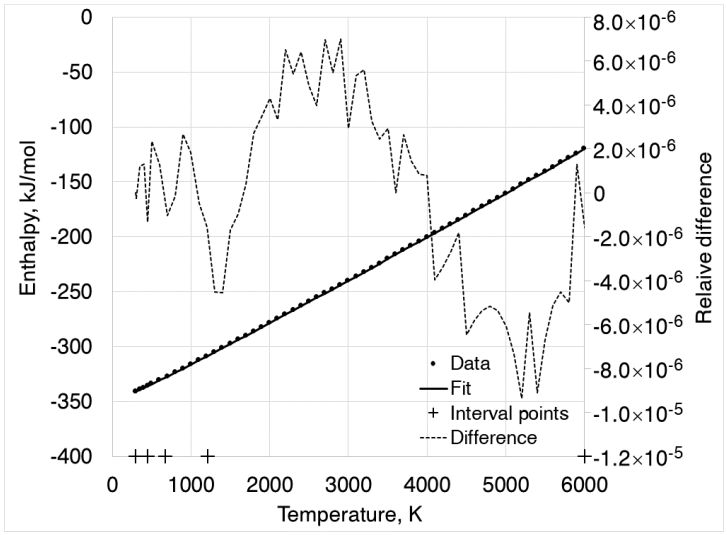


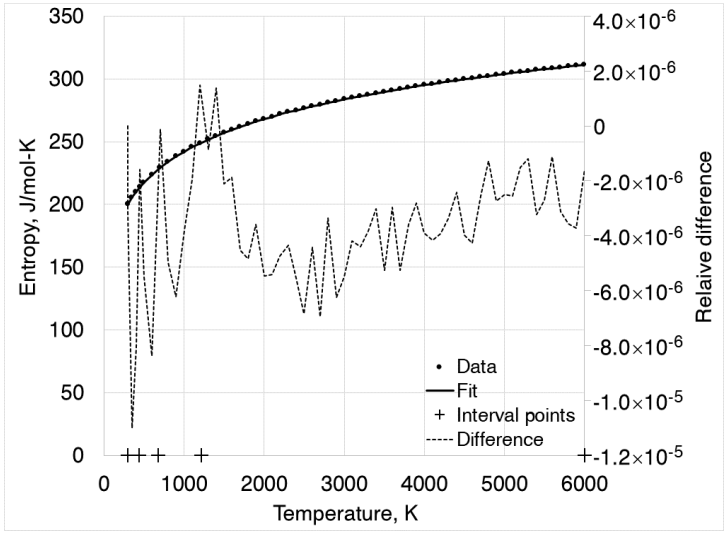


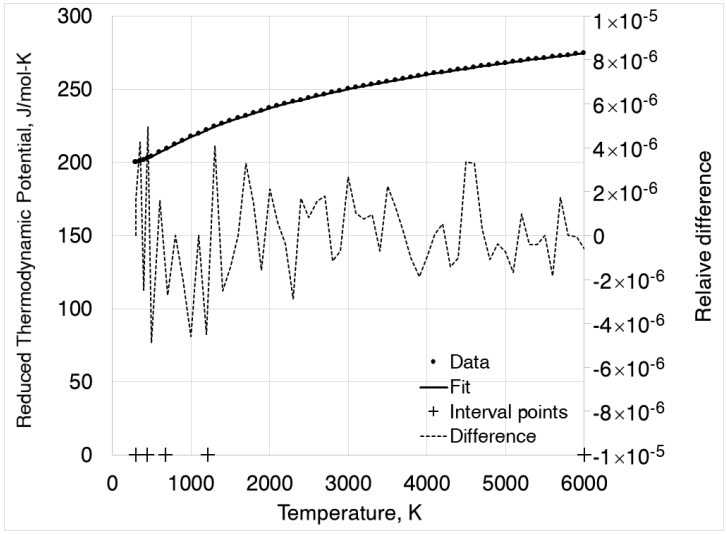
Lithium Fluoride (LiF) – F1Li1(g)

MW = 25.939 g/mol – $\Delta_f H^o$(298.15 K) = -340.787 kJ/mol – S^o(298.15 K) = 200.278 J/mol-K – p^o = 0.1 MPa



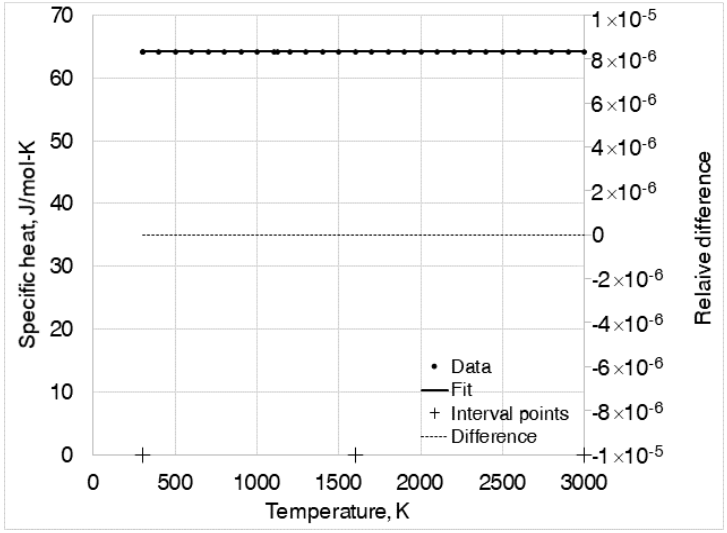


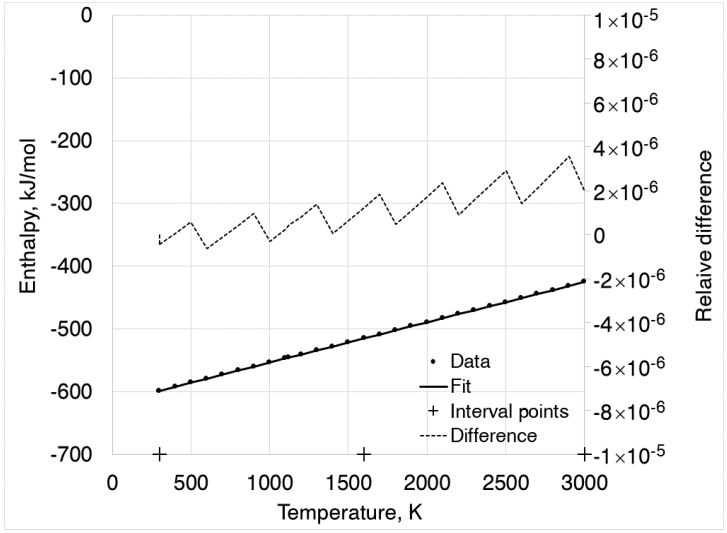


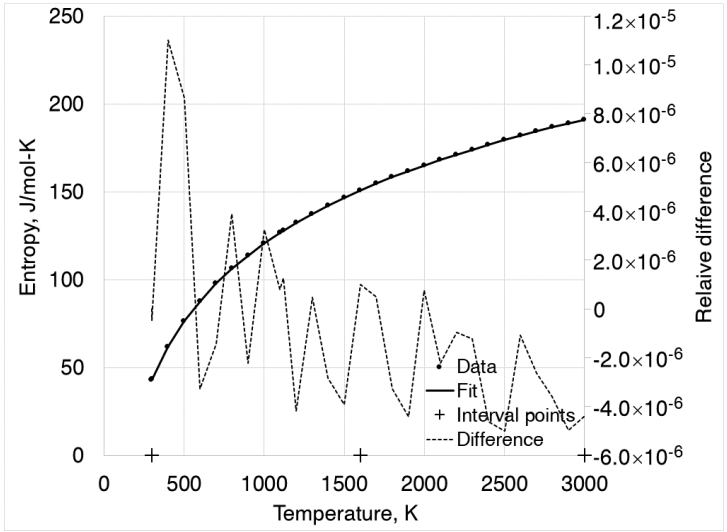


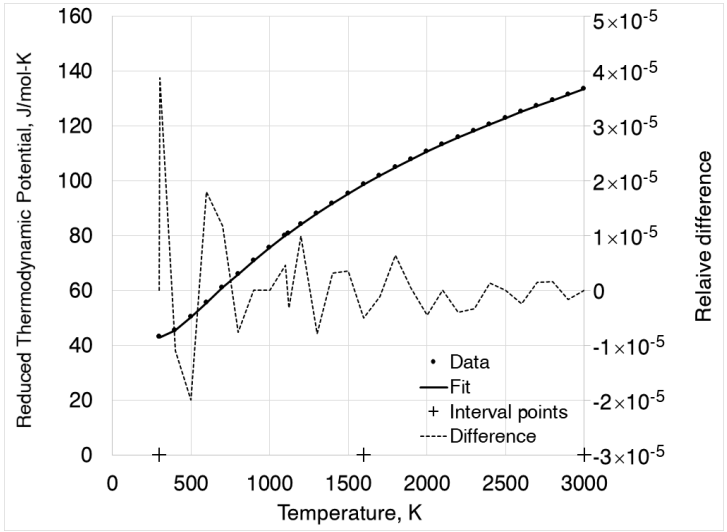
Lithium Fluoride (LiF) – F1Li1(l)

MW = 25.939 g/mol – $\Delta_f H^o$(298.15 K) = -598.651 kJ/mol – S^o(298.15 K) = 42.962 J/mol-K – p^o = 0.1 MPa



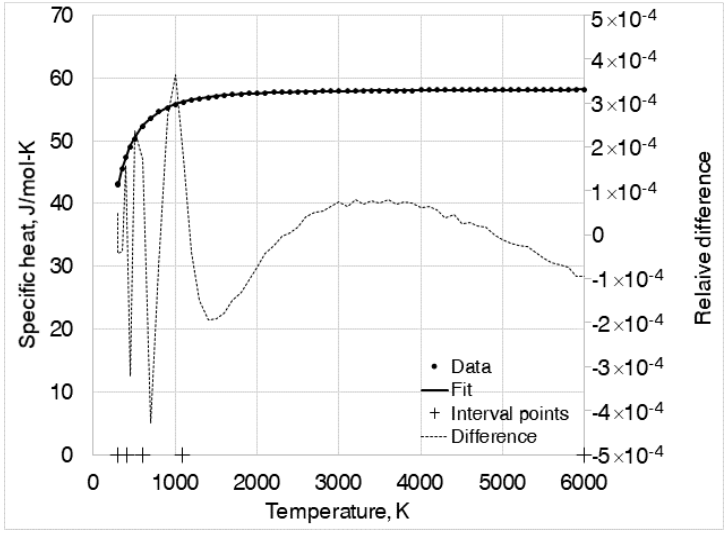


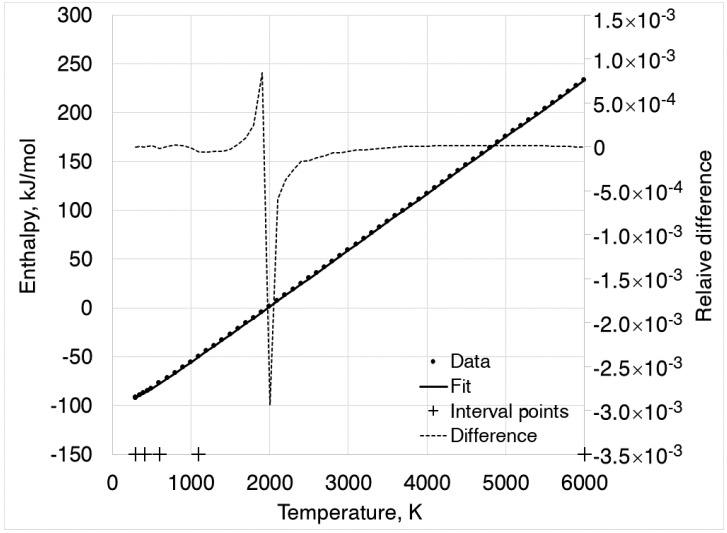


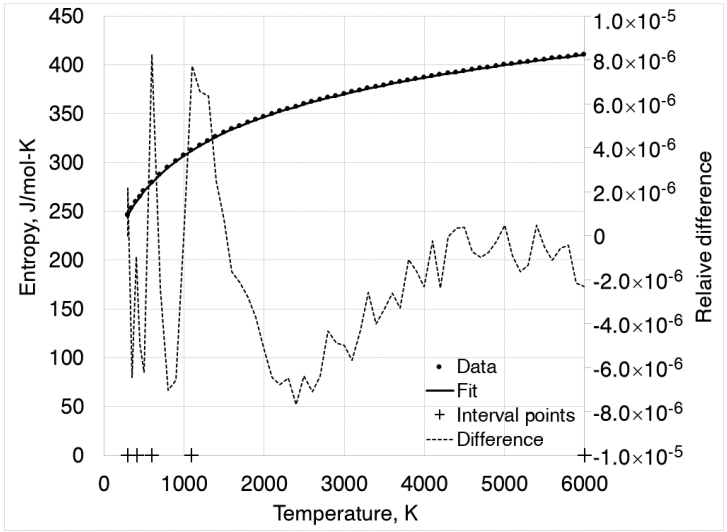


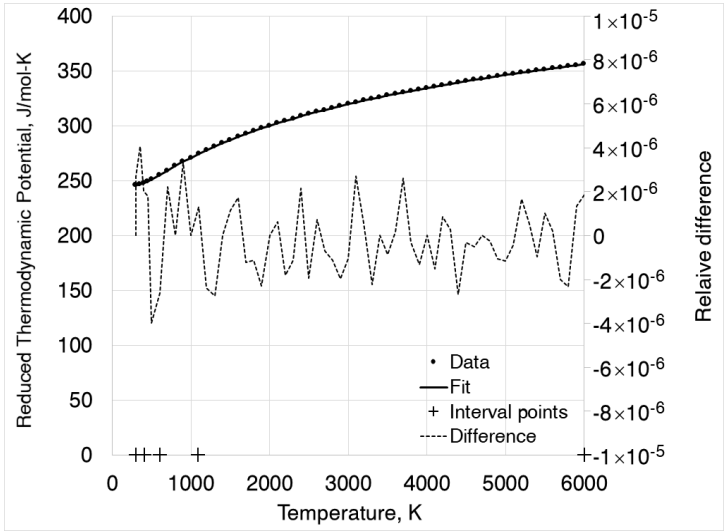
Lithium Hypofluorite (LiOF) – F1Li1O1(g)

MW = 41.939 g/mol – $\Delta_f H^o$(298.15 K) = -92.048 kJ/mol – S^o(298.15 K) = 246.022 J/mol-K – p^o = 0.1 MPa



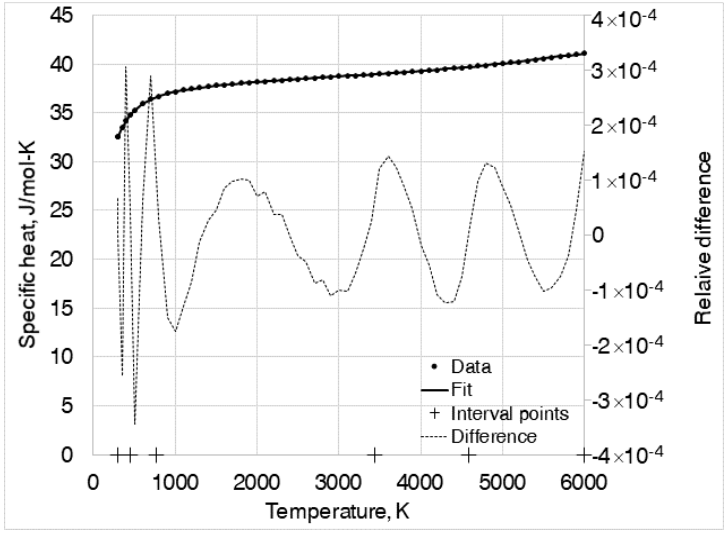


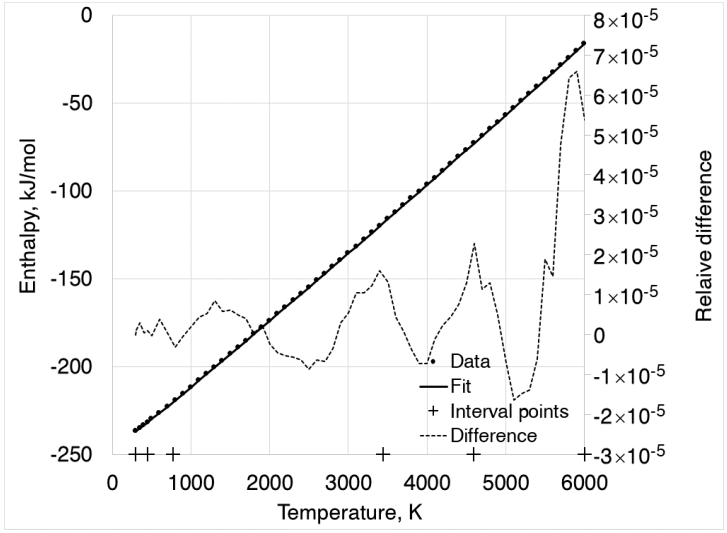


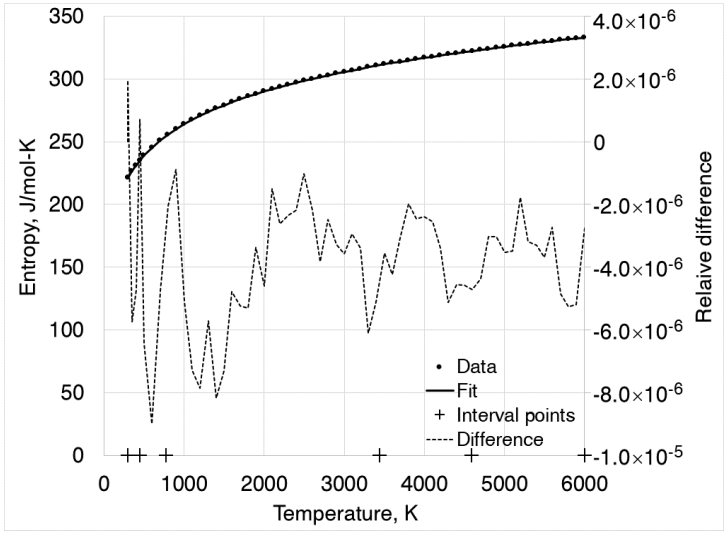


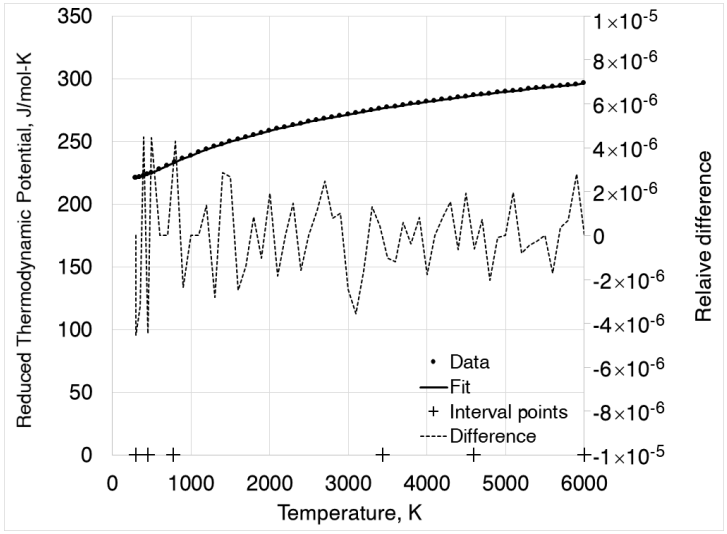
Magnesium Fluoride (MgF) – F1Mg1(g)

MW = 43.303 g/mol – $\Delta_f H^o$(298.15 K) = -236.814 kJ/mol – S^o(298.15 K) = 221.085 J/mol-K – p^o = 0.1 MPa



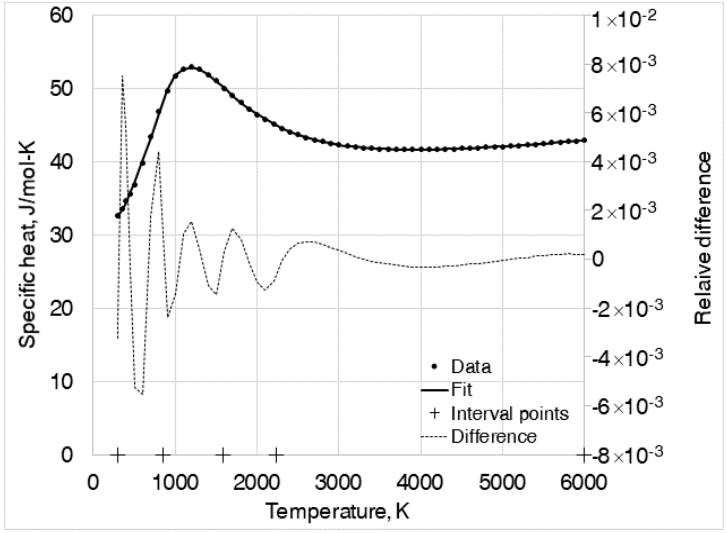


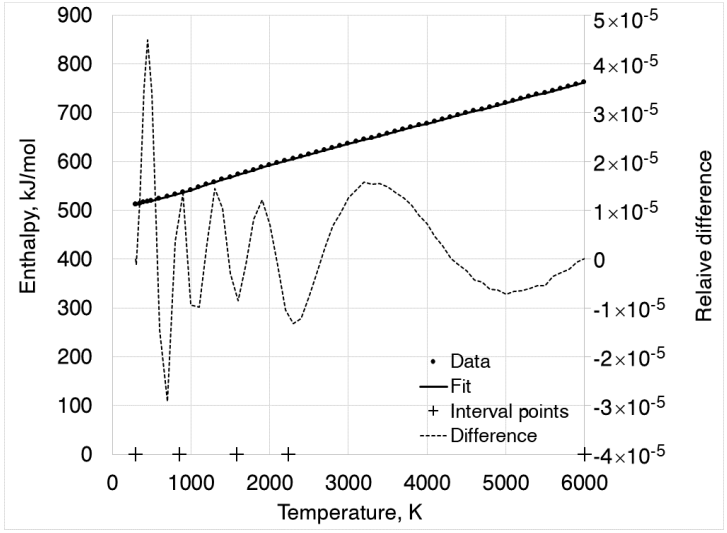


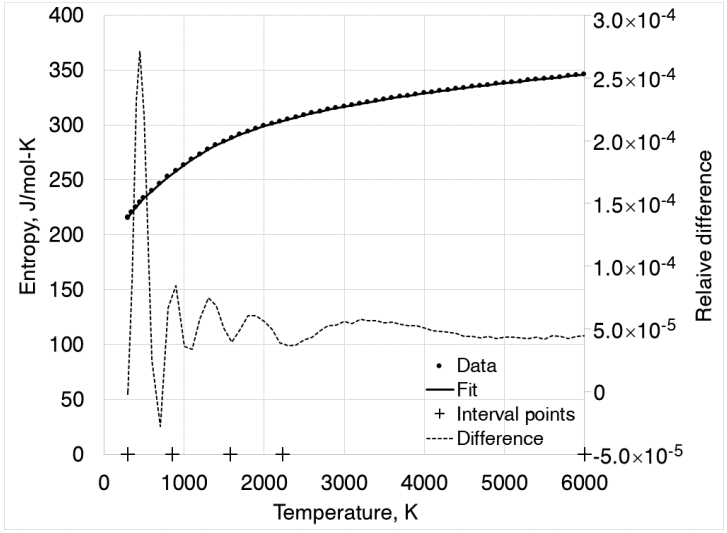


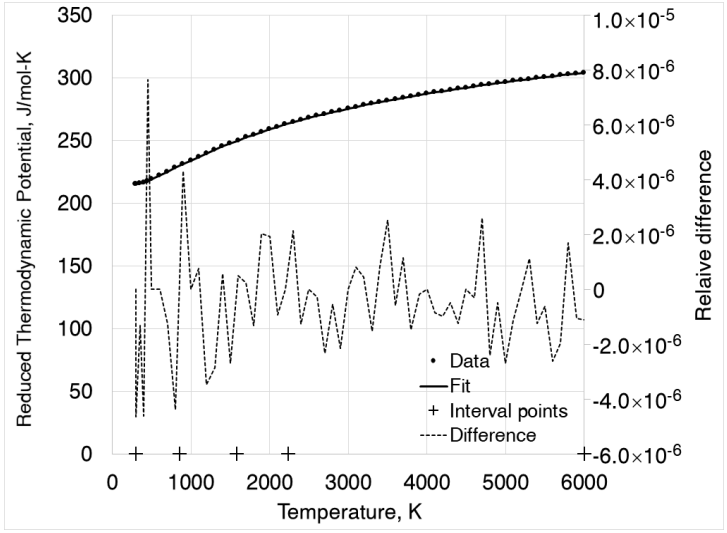
Magnesium Fluoride, Ion (MgF^+) – F1Mg1+(g)

MW = 43.303 g/mol – $\Delta_f H^o$(298.15 K) = 512.289 kJ/mol – S^o(298.15 K) = 215.33 J/mol-K – p^o = 0.1 MPa



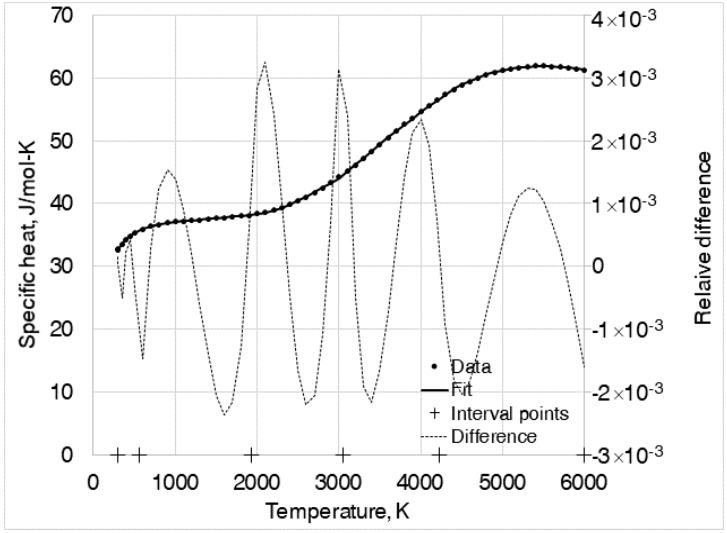


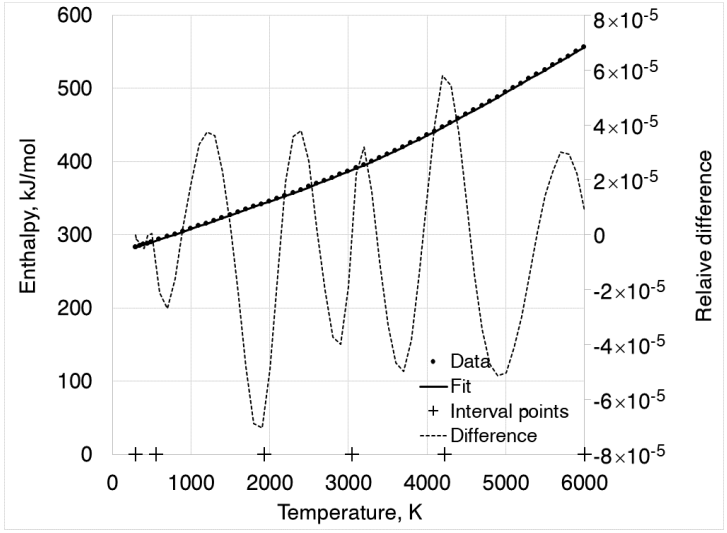


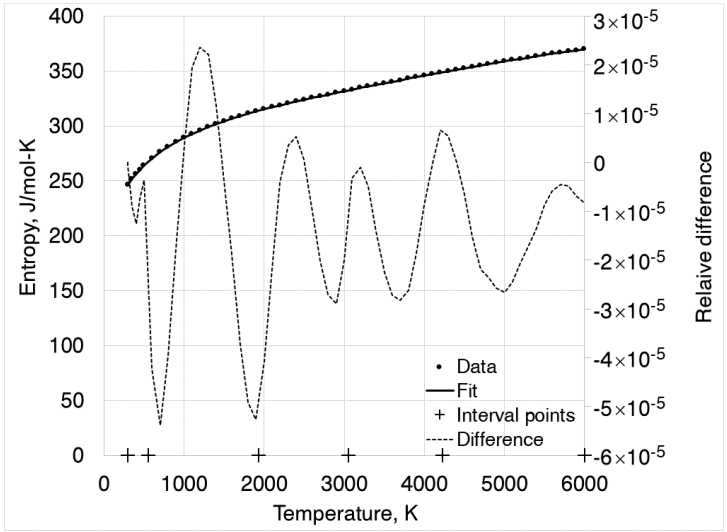


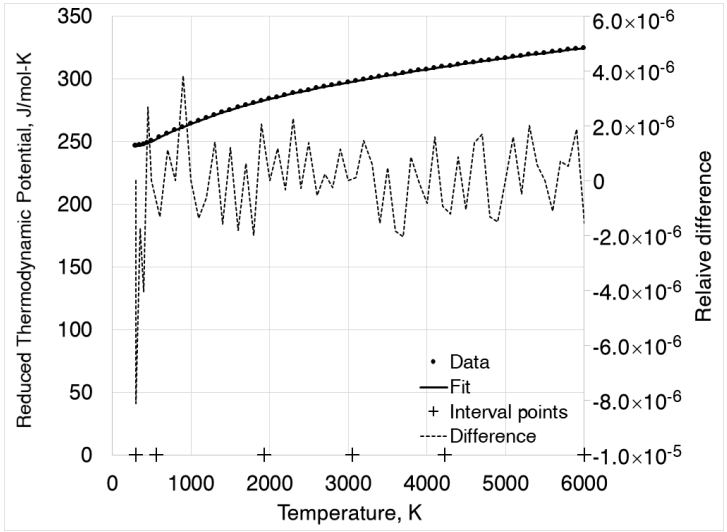
Molybdenum Fluoride (MoF) – F1Mo1(g)

MW = 114.94 g/mol – $\Delta_f H^o$(298.15 K) = 282.838 kJ/mol – S^o(298.15 K) = 246.399 J/mol-K – p^o = 0.1 MPa



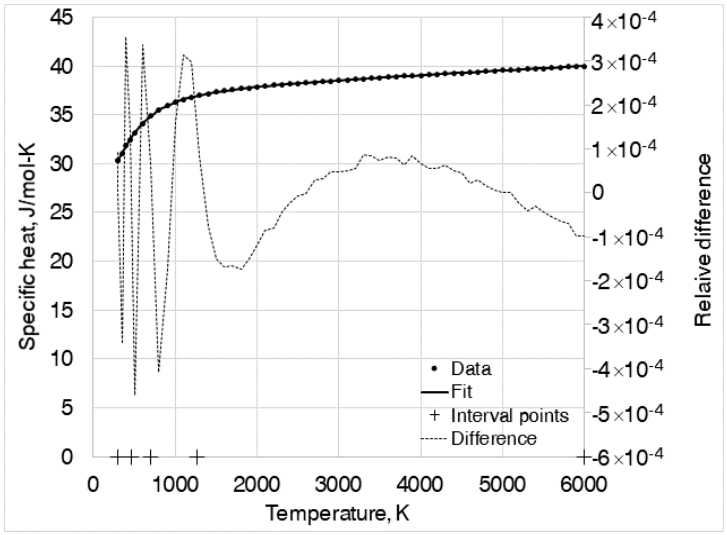


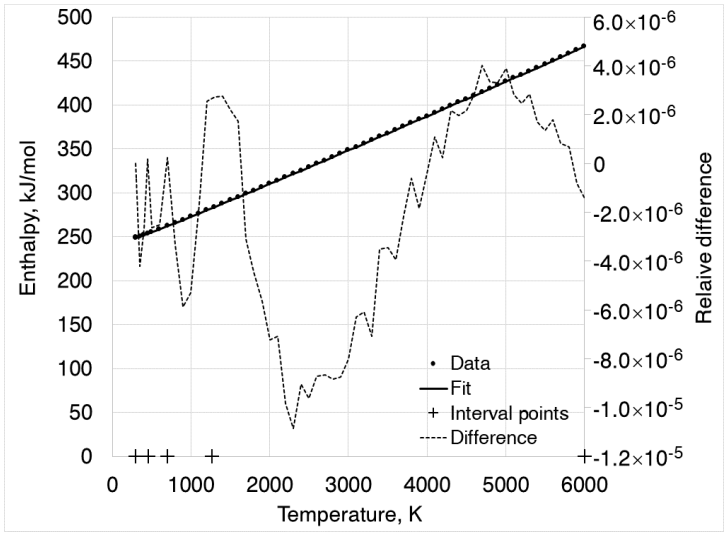


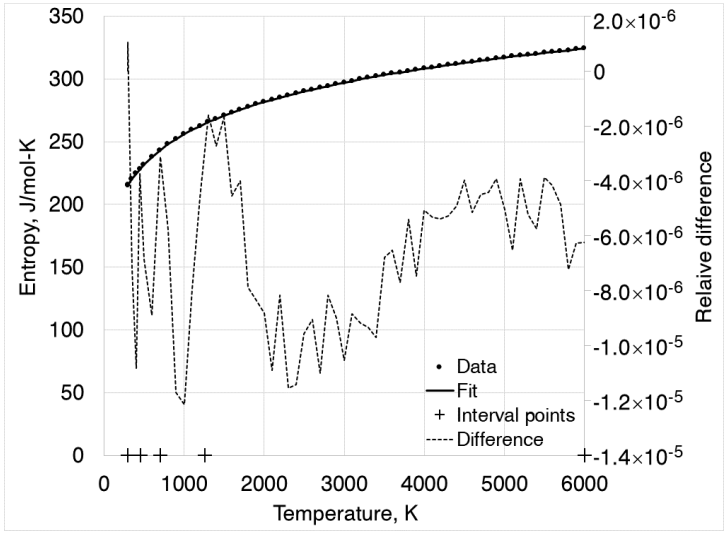


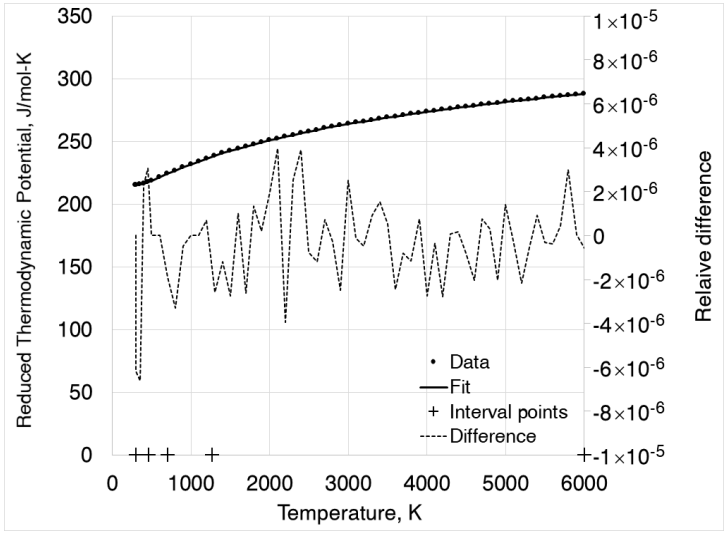
Fluoroimidogen (NF) – F1N1(g)

MW = 33.0051 g/mol – $\Delta_f H^o$(298.15 K) = 248.948 kJ/mol – S^o(298.15 K) = 215.276 J/mol-K – p^o = 0.1 MPa



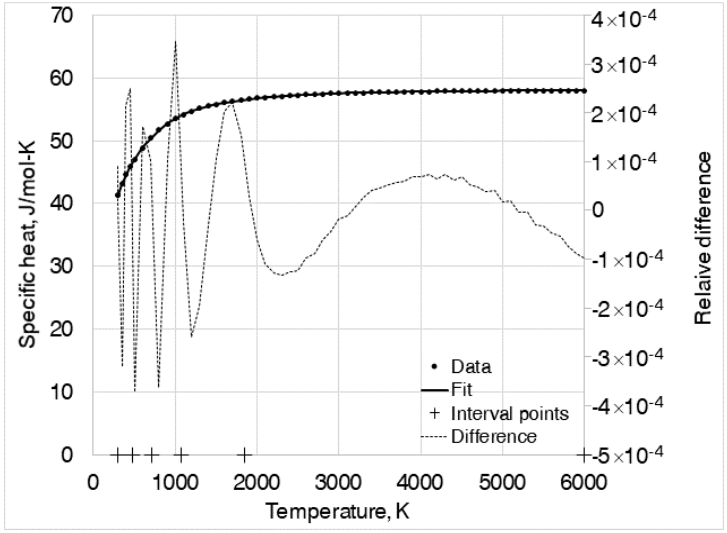


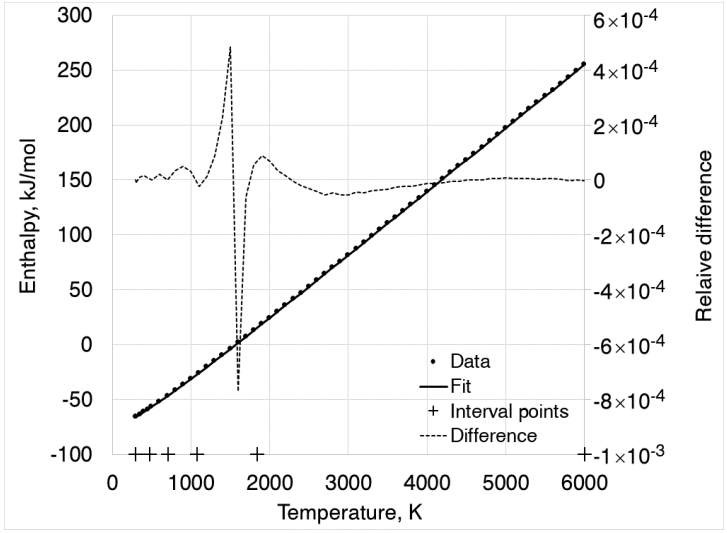


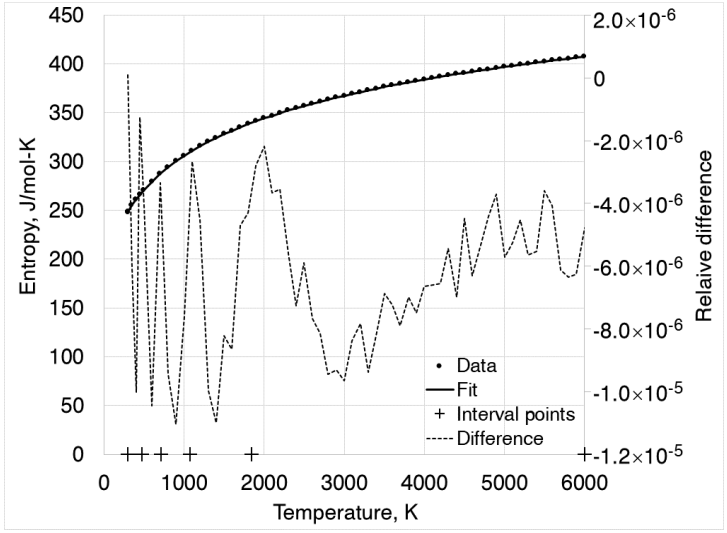


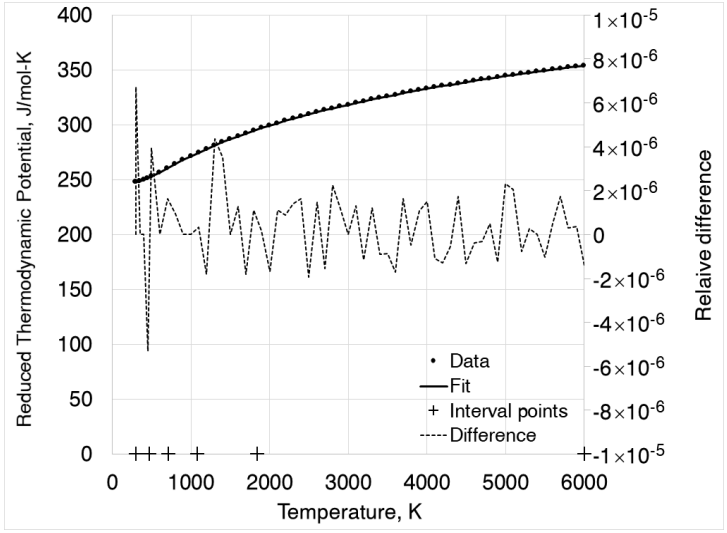
Nitrosyl Fluoride (ONF) – F1N1O1(g)

MW = 49.0045 g/mol – $\Delta_f H^o$(298.15 K) = -65.689 kJ/mol – S^o(298.15 K) = 248.082 J/mol-K – p^o = 0.1 MPa



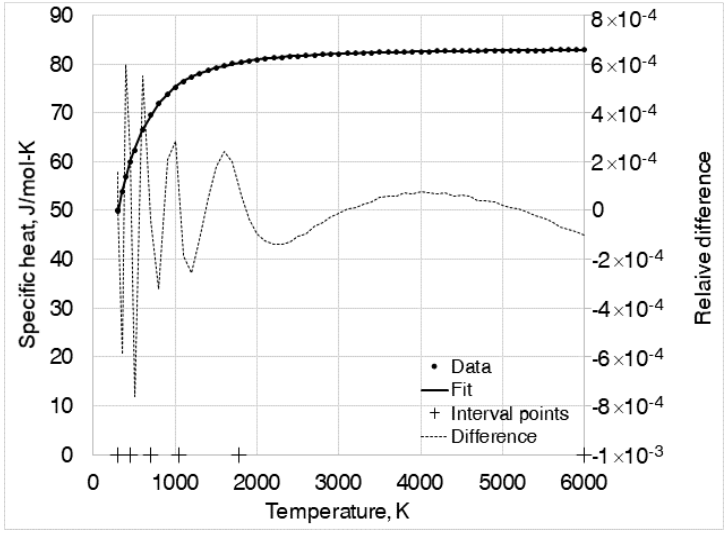


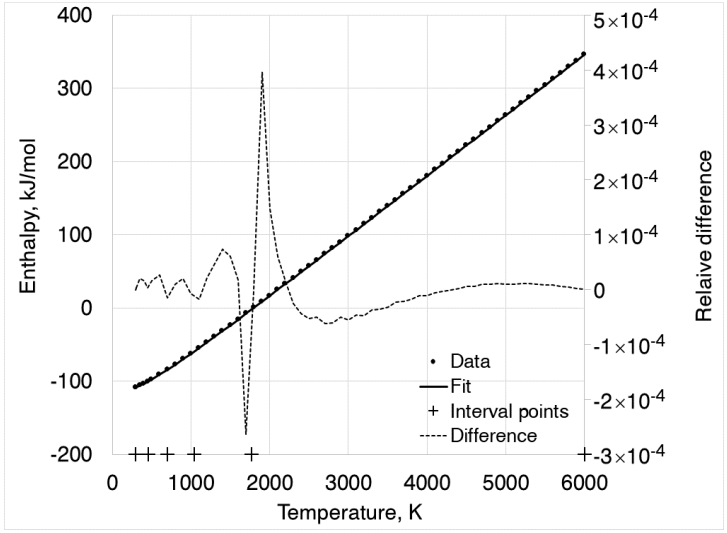


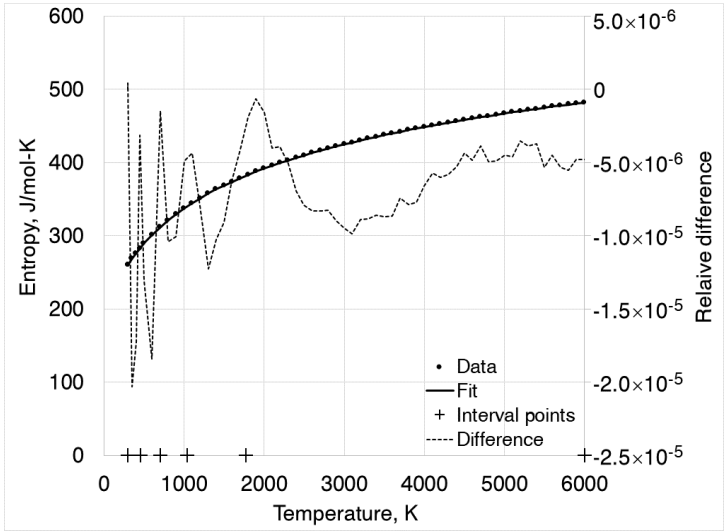


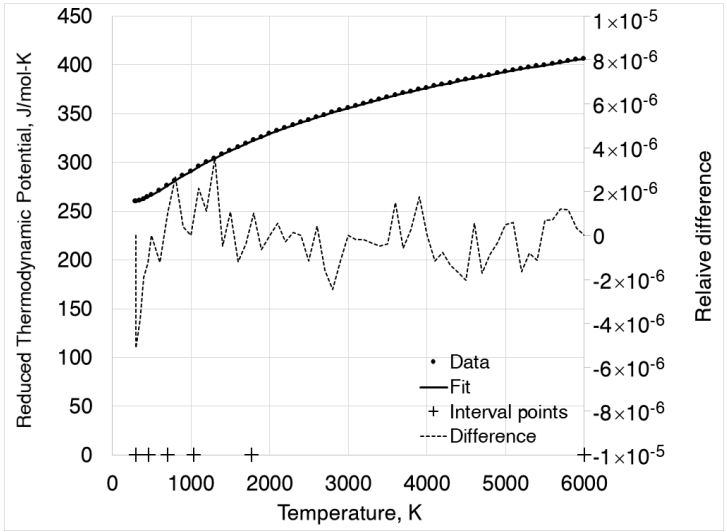
Nitryl Fluoride (NO_2F) – F1N1O2(g)

MW = 65.0039 g/mol – $\Delta_f H^o$(298.15 K) = -108.784 kJ/mol – S^o(298.15 K) = 260.25 J/mol-K – p^o = 0.1 MPa



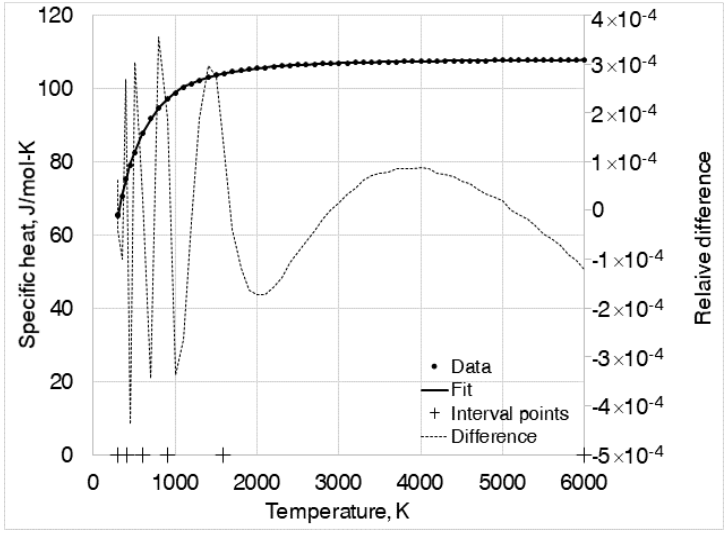


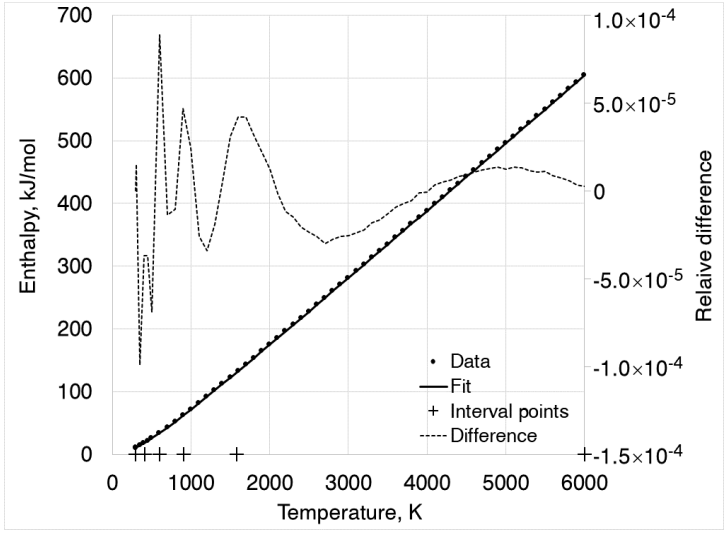


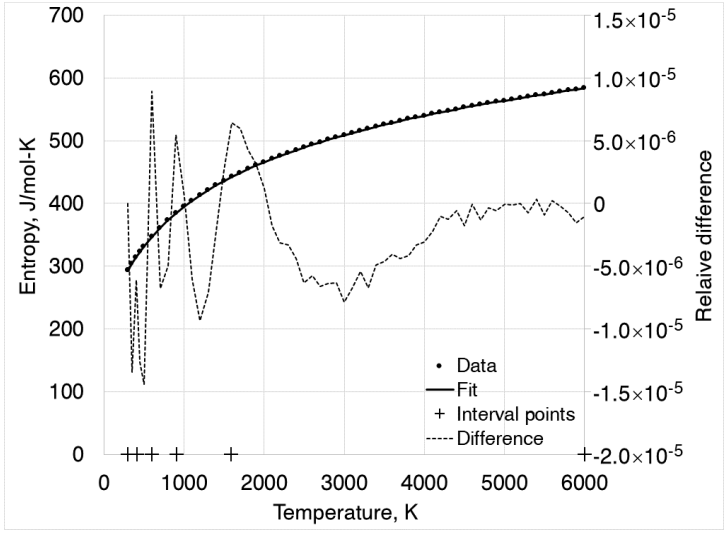


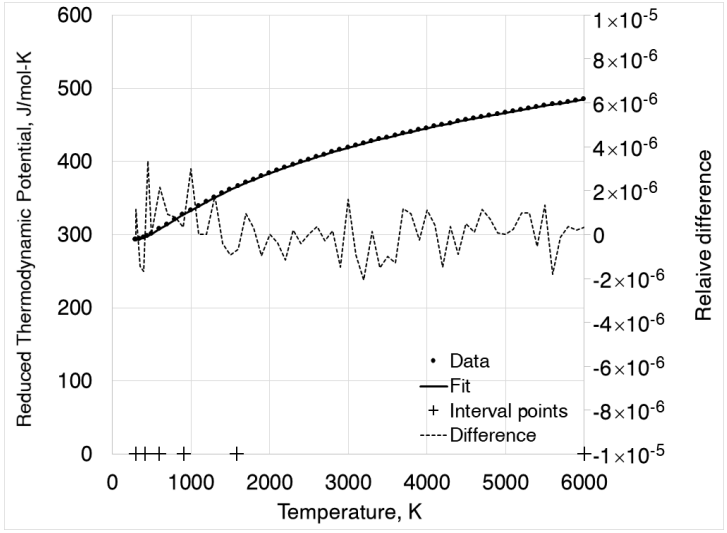
Fluorine Nitrate ($FONO_2$) – F1N1O3(g)

MW = 81.0033 g/mol – $\Delta_f H^o$(298.15 K) = 10.46 kJ/mol – S^o(298.15 K) = 292.933 J/mol-K – p^o = 0.1 MPa



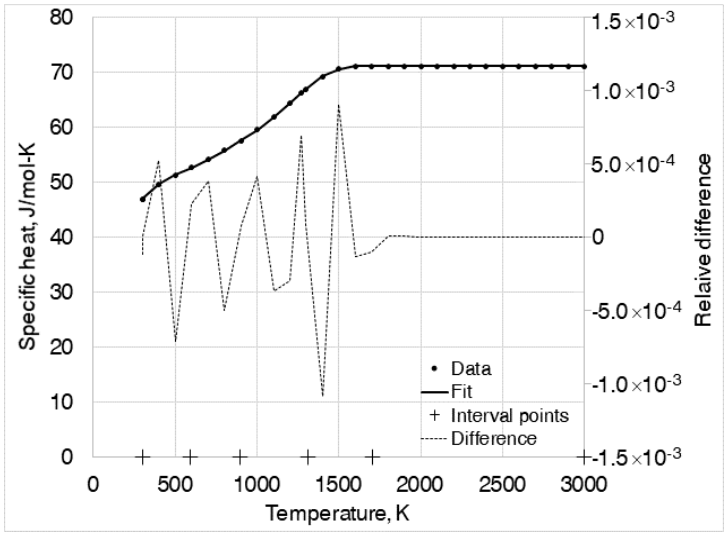


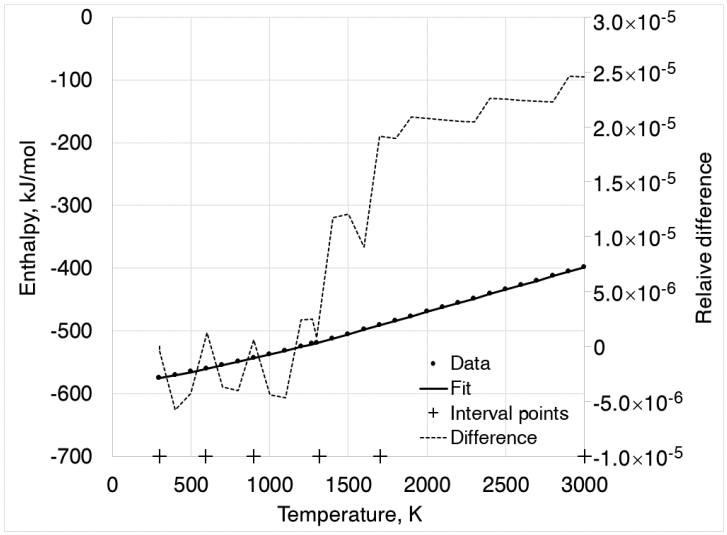


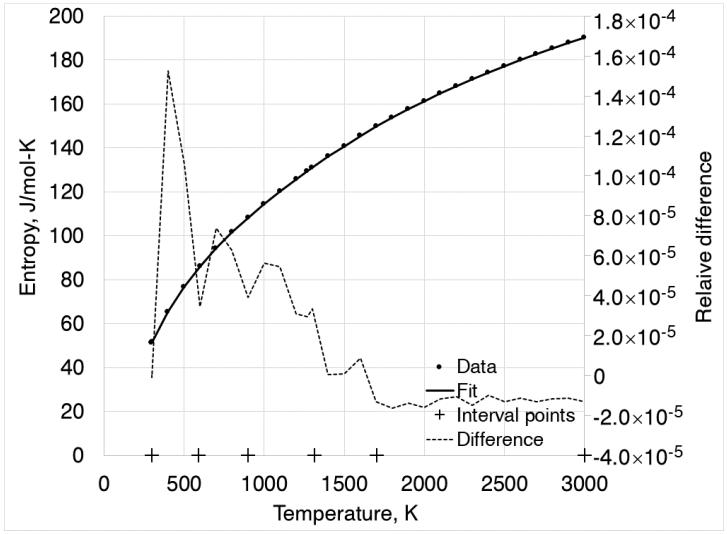


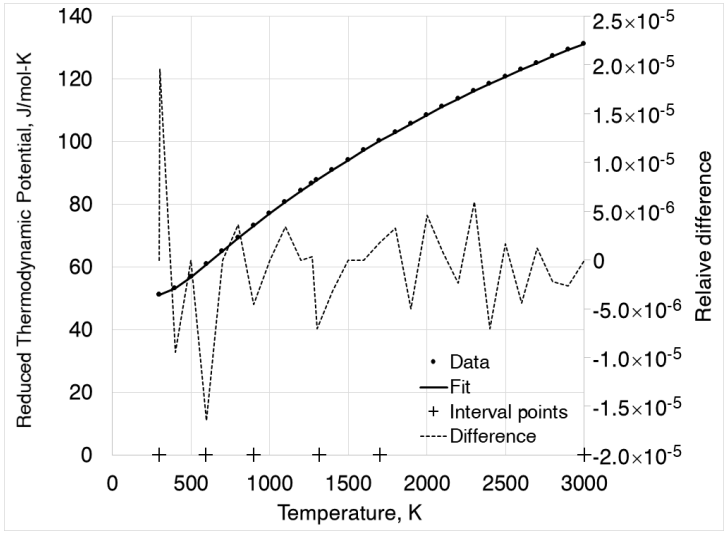
Sodium Fluoride (NaF) – F1Na1(cr)

MW = 41.9882 g/mol – $\Delta_f H^o$(298.15 K) = -575.384 kJ/mol – S^o(298.15 K) = 51.212 J/mol-K – p^o = 0.1 MPa



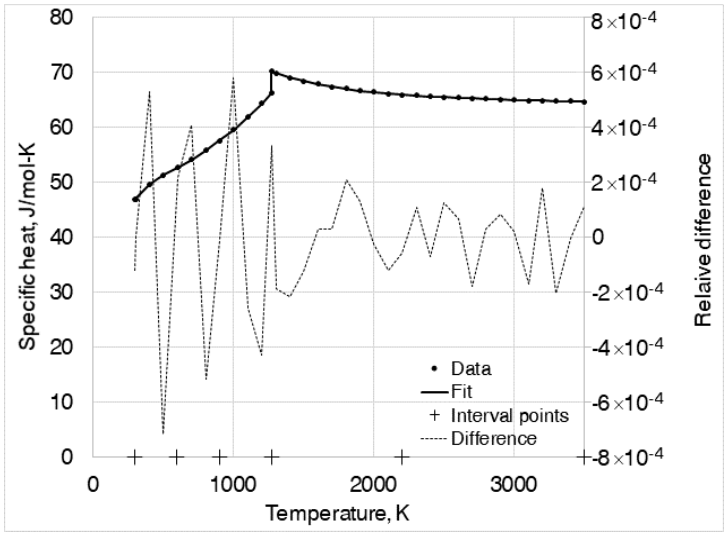


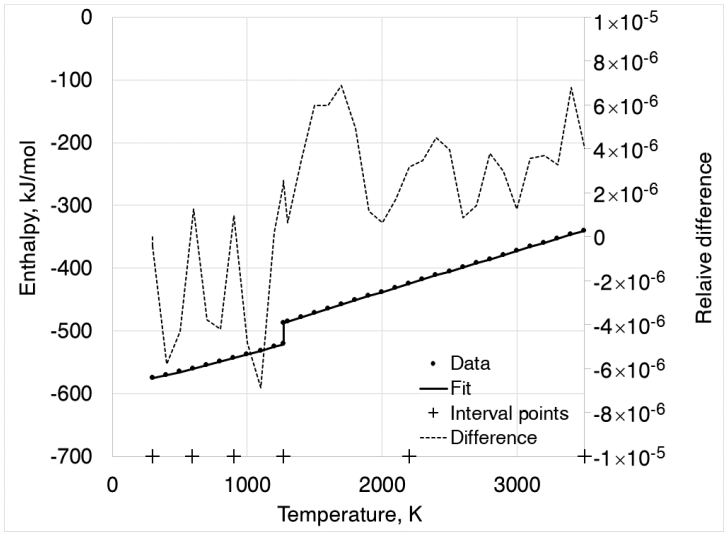


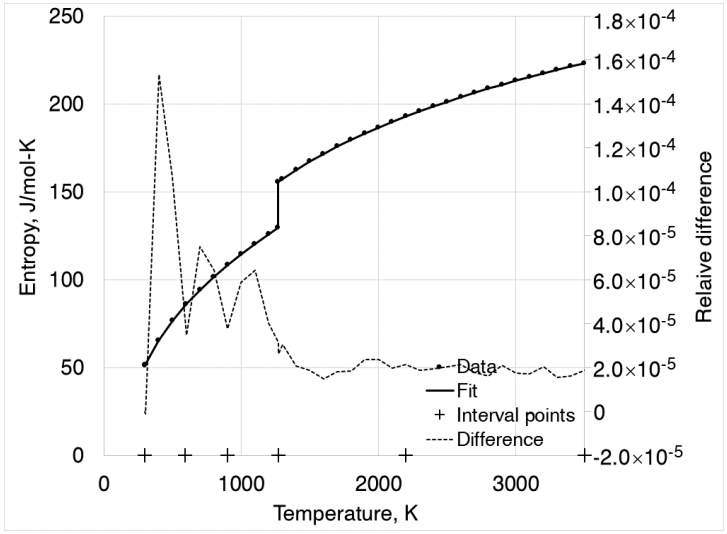


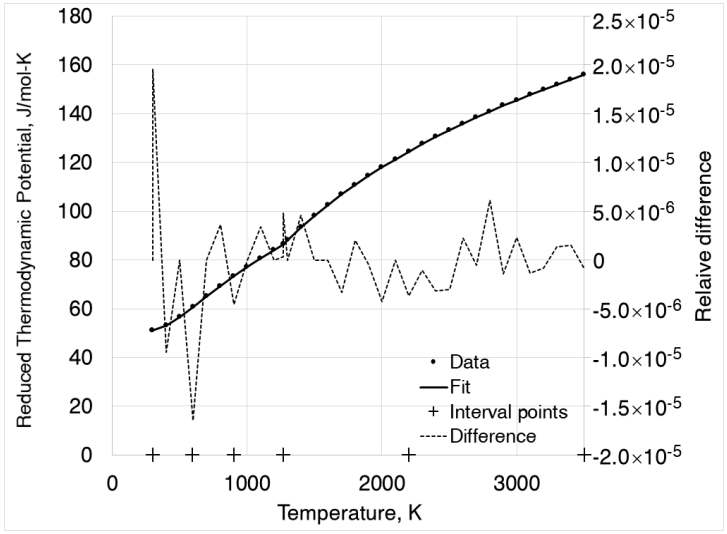
Sodium Fluoride (NaF) – F1Na1(cr,l)

MW = 41.9882 g/mol – $\Delta_f H^o$(298.15 K) = -575.384 kJ/mol – S^o(298.15 K) = 51.212 J/mol-K – p^o = 0.1 MPa



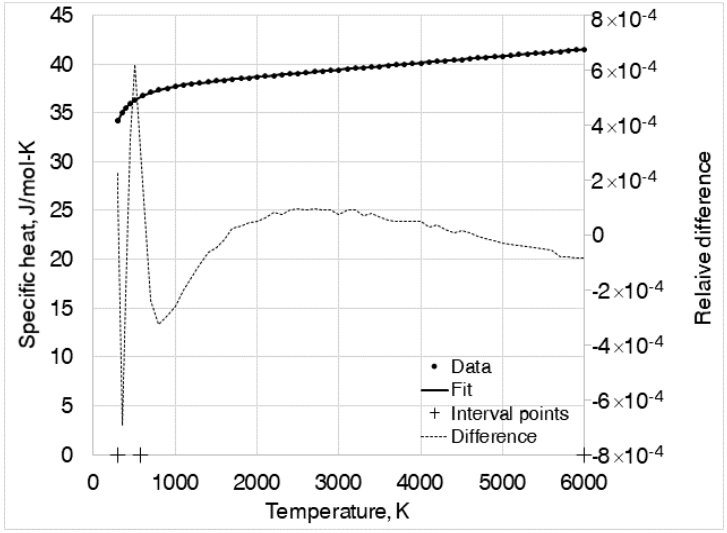


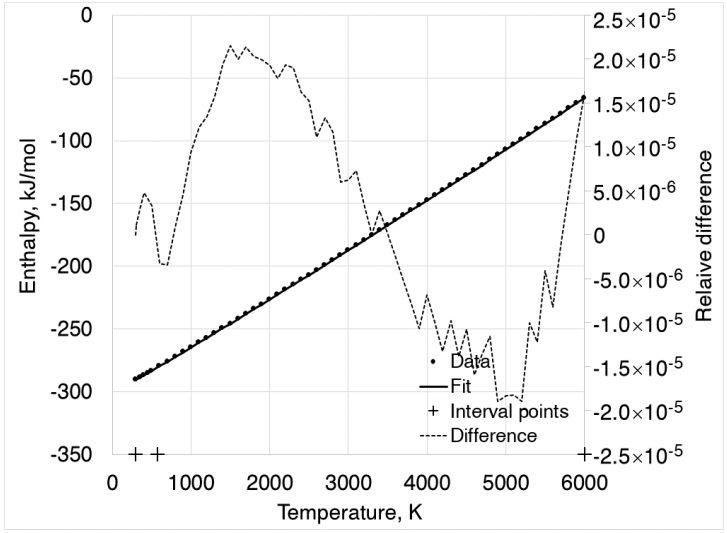


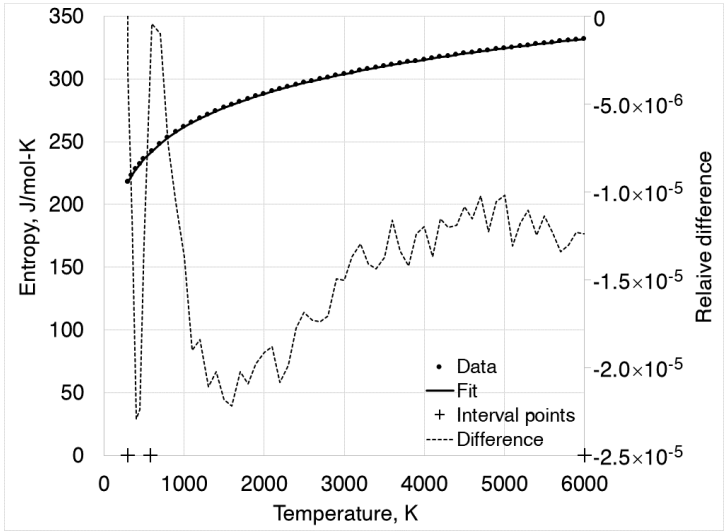


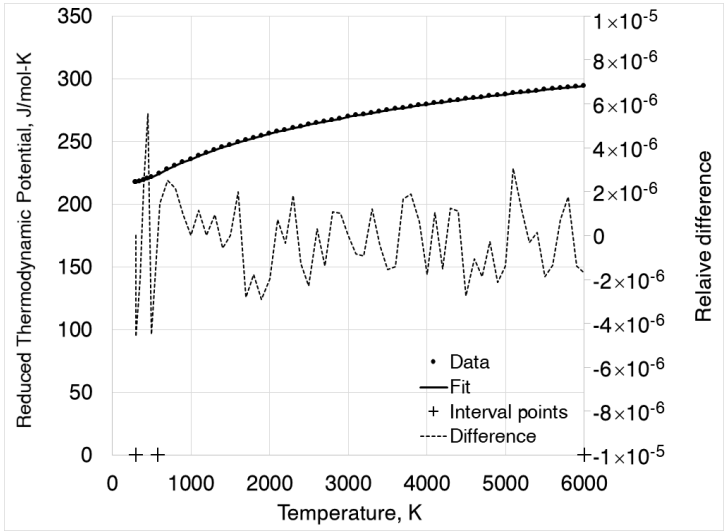
Sodium Fluoride (NaF) – F1Na1(g)

MW = 41.9882 g/mol – $\Delta_f H^o$(298.15 K) = -290.453 kJ/mol – S^o(298.15 K) = 217.61 J/mol-K – p^o = 0.1 MPa



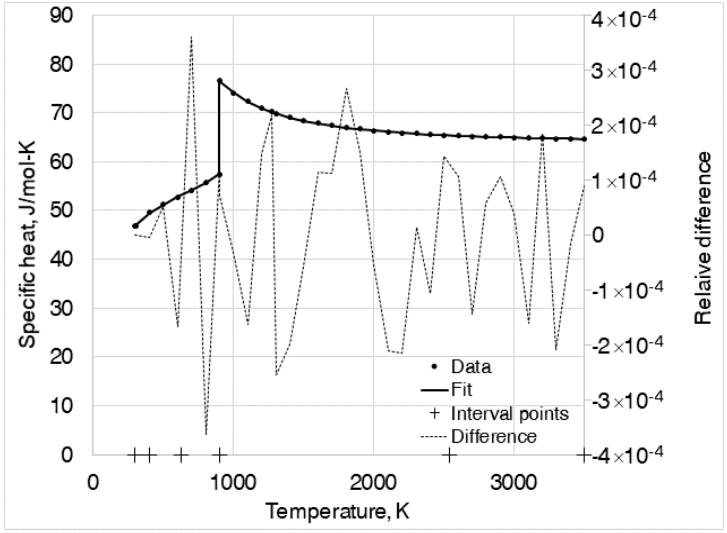


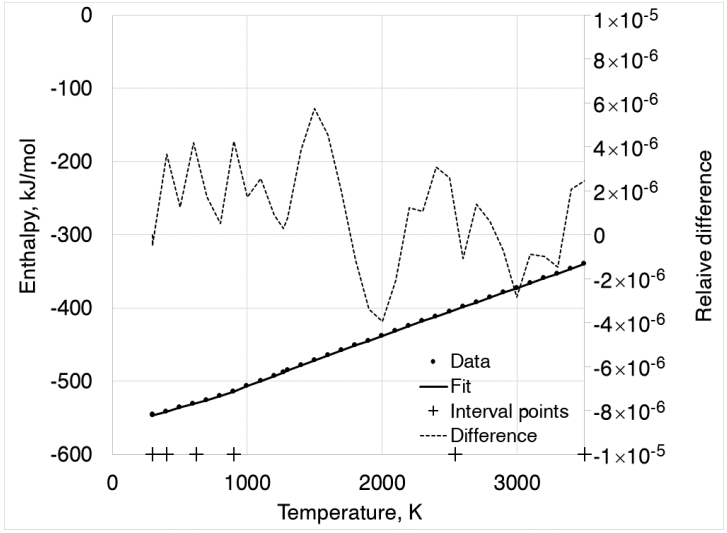


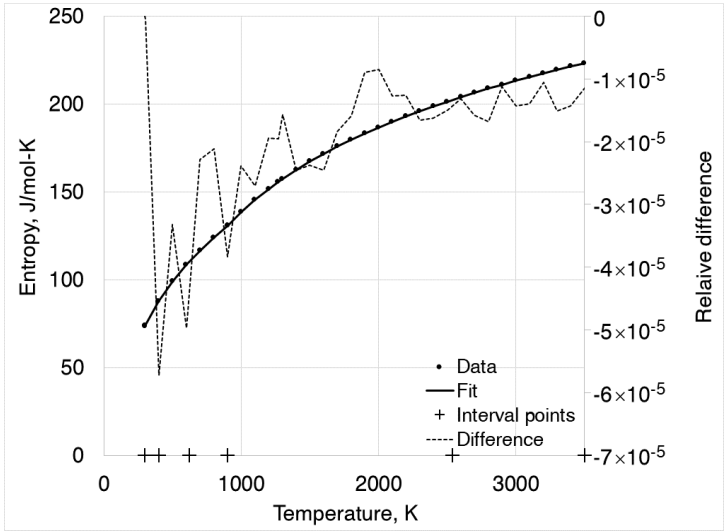


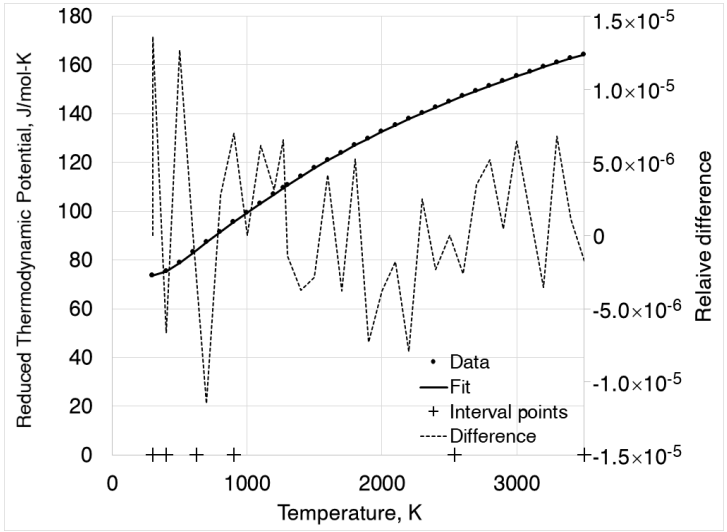
Sodium Fluoride (NaF) – F1Na1(l)

MW = 41.9882 g/mol – $\Delta_f H^o$(298.15 K) = -546.204 kJ/mol – S^o(298.15 K) = 73.464 J/mol-K – p^o = 0.1 MPa



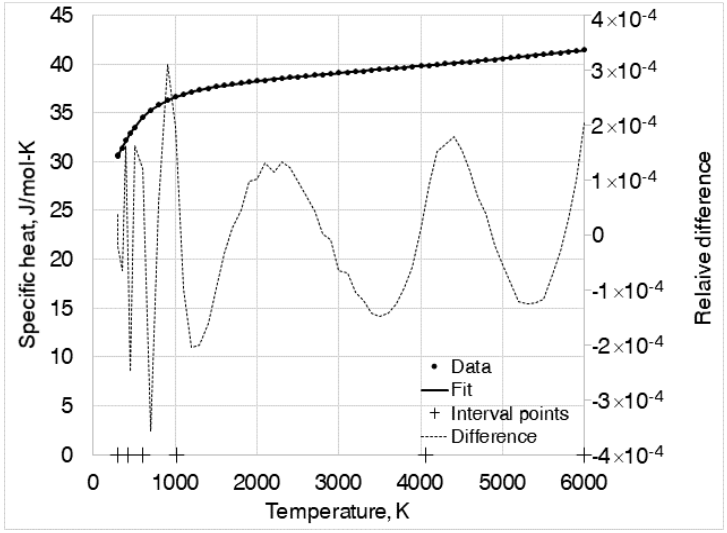


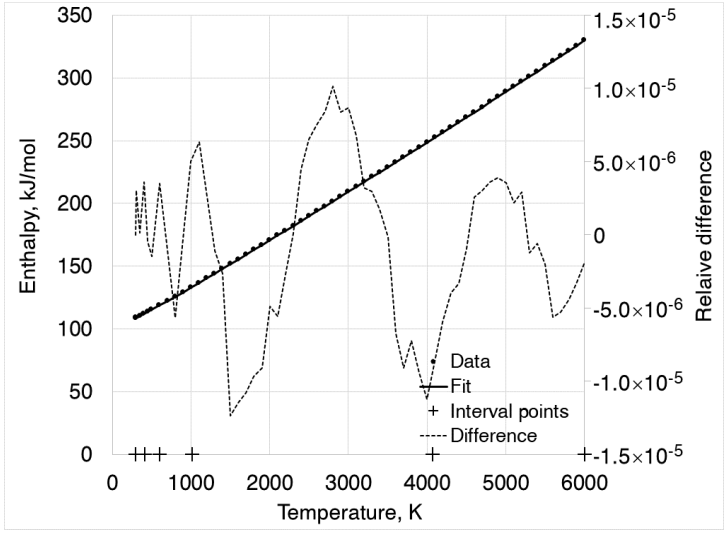


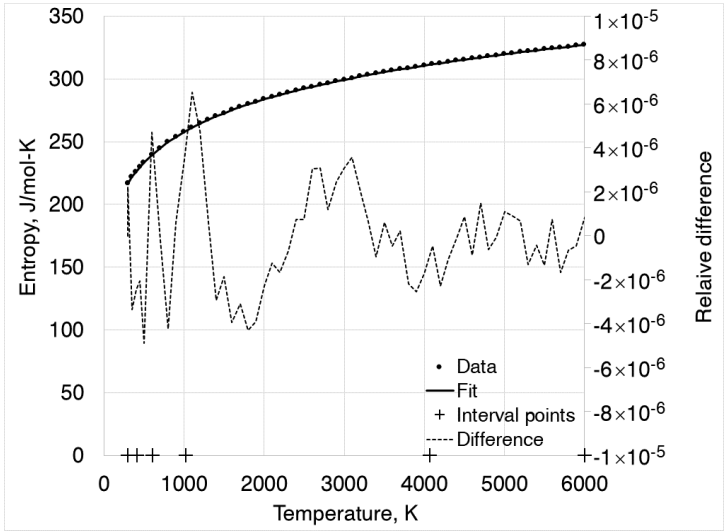


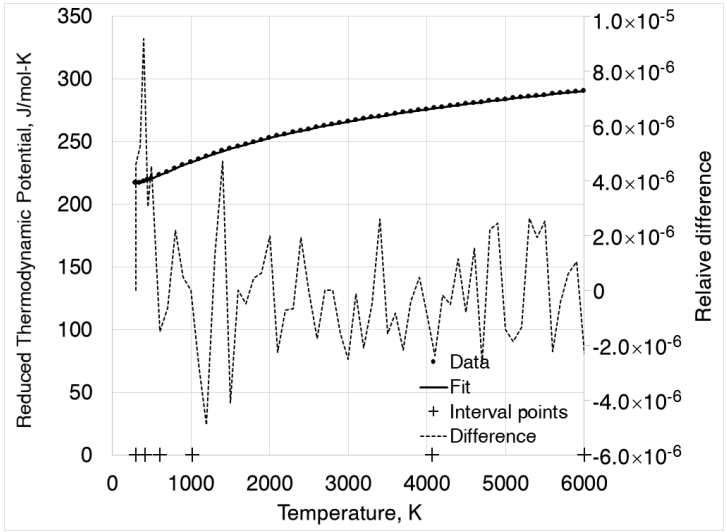
Oxygen Fluoride (OF) – F1O1(g)

MW = 34.9978 g/mol – $\Delta_f H^o$(298.15 K) = 108.784 kJ/mol – S^o(298.15 K) = 216.696 J/mol-K – p^o = 0.1 MPa



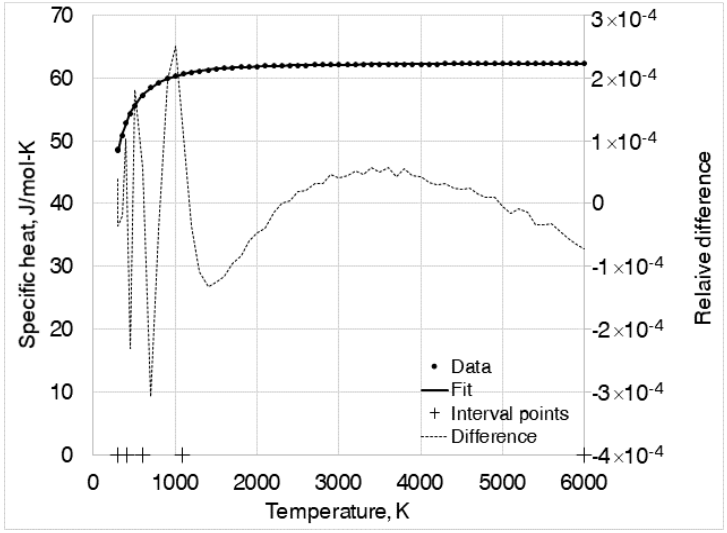


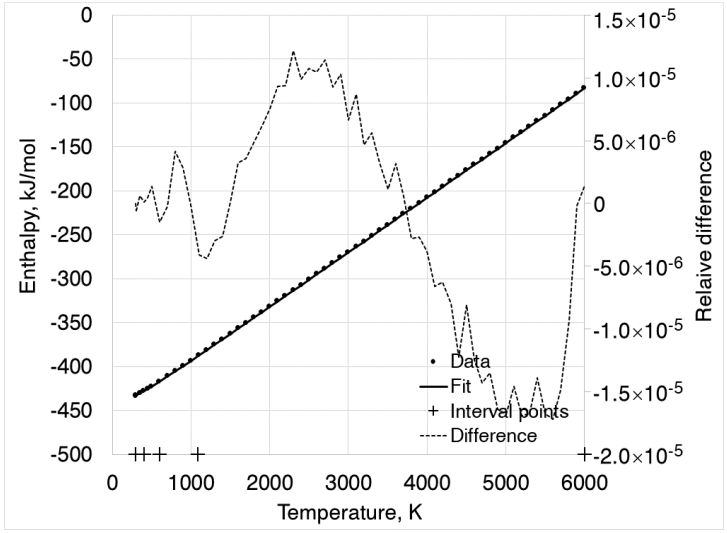


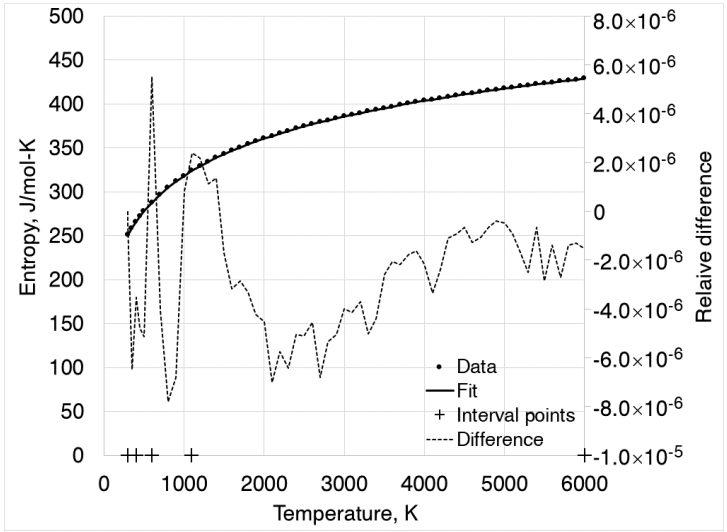


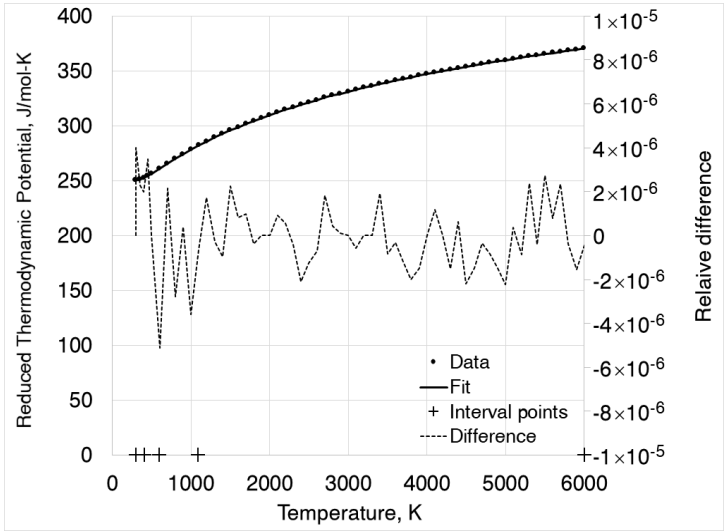
Titanium Fluoride Oxide (OTiF) – F1O1Ti1(g)

MW = 82.88 g/mol – $\Delta_f H^o$(298.15 K) = -433.211 kJ/mol – S^o(298.15 K) = 250.667 J/mol-K – p^o = 0.1 MPa



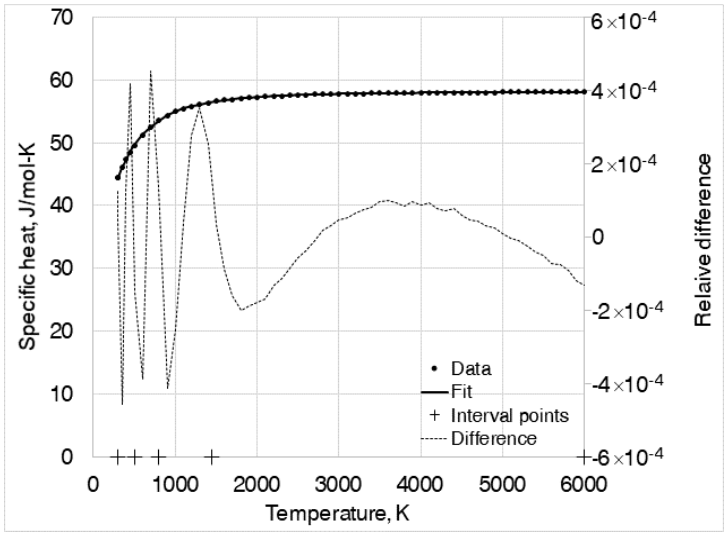


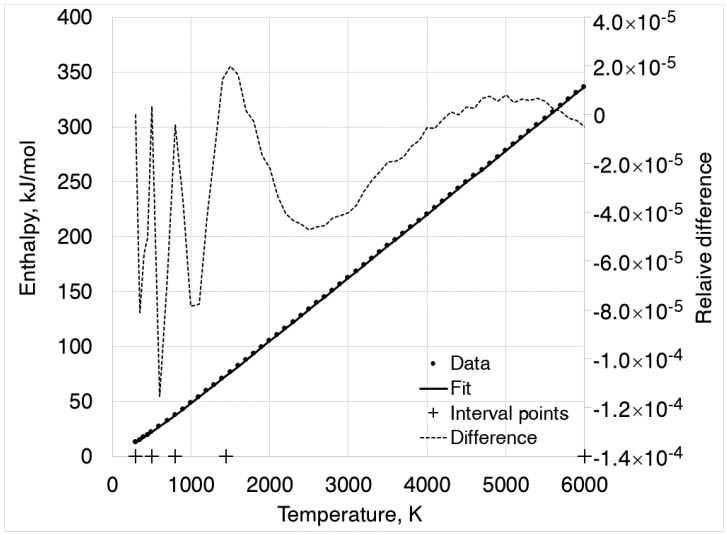


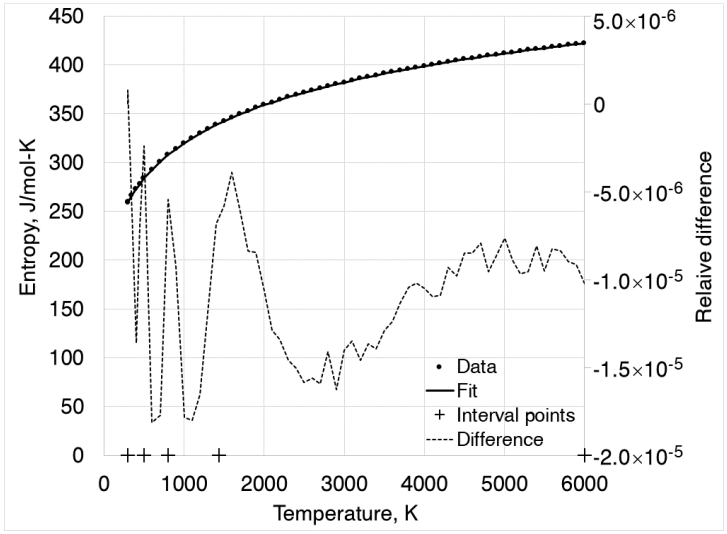


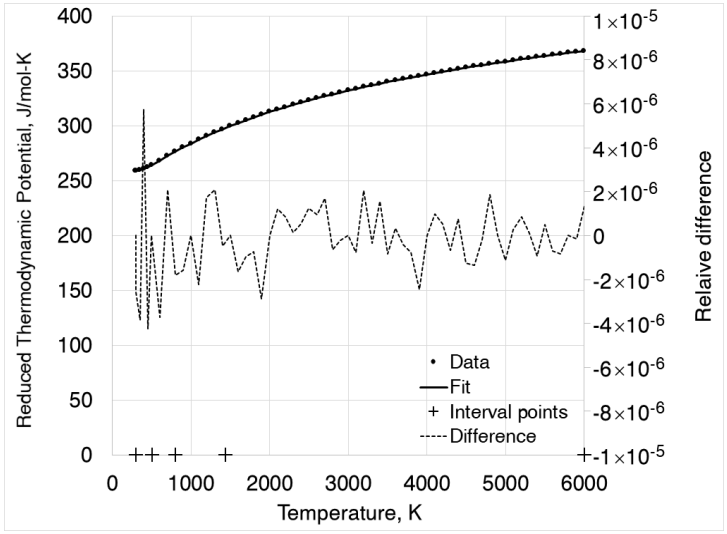
Oxygen Fluoride (O_2F) – F1O2(g)

MW = 50.9972 g/mol – $\Delta_f H^o$(298.15 K) = 12.552 kJ/mol – S^o(298.15 K) = 259.114 J/mol-K – p^o = 0.1 MPa



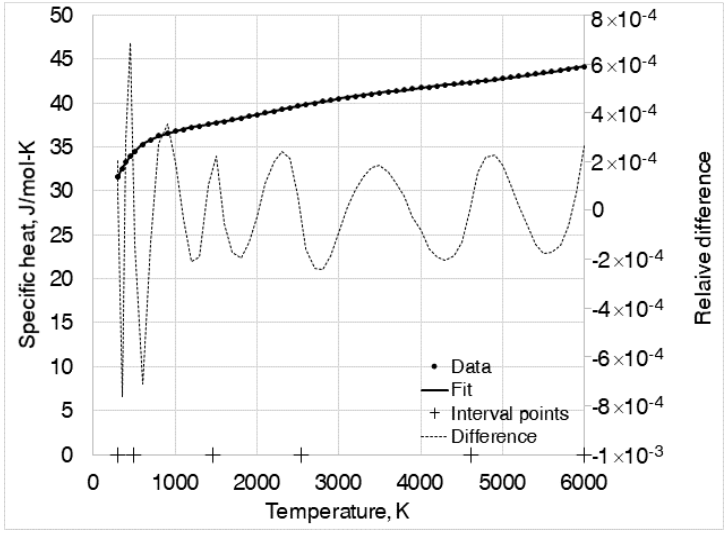


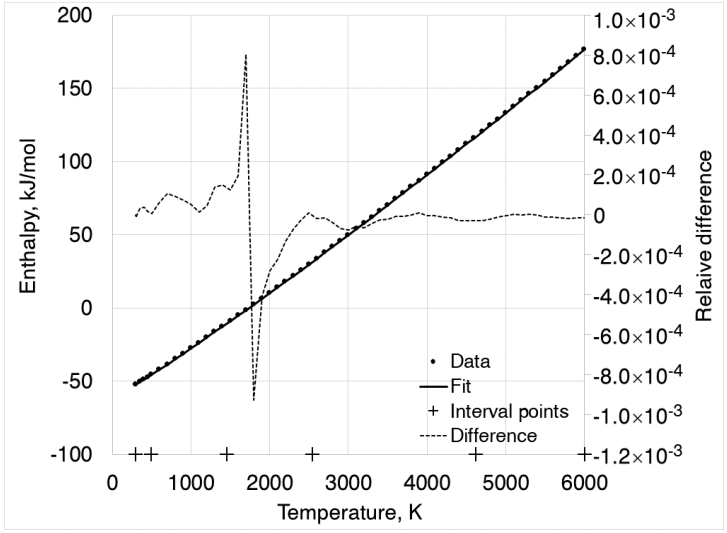


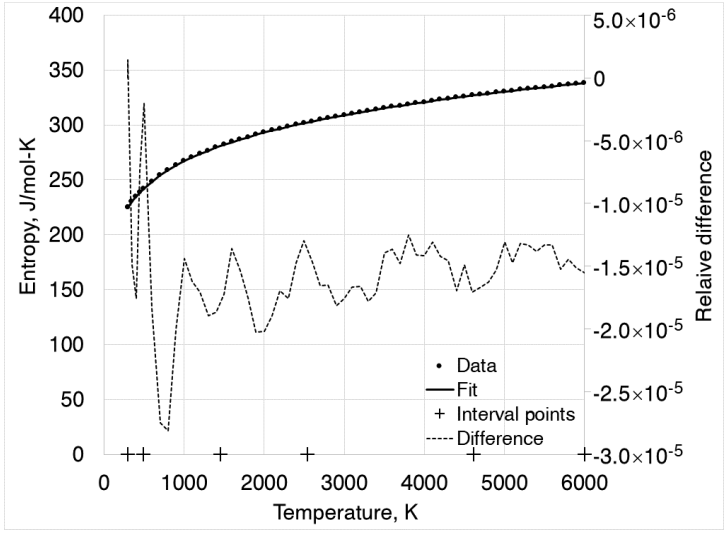


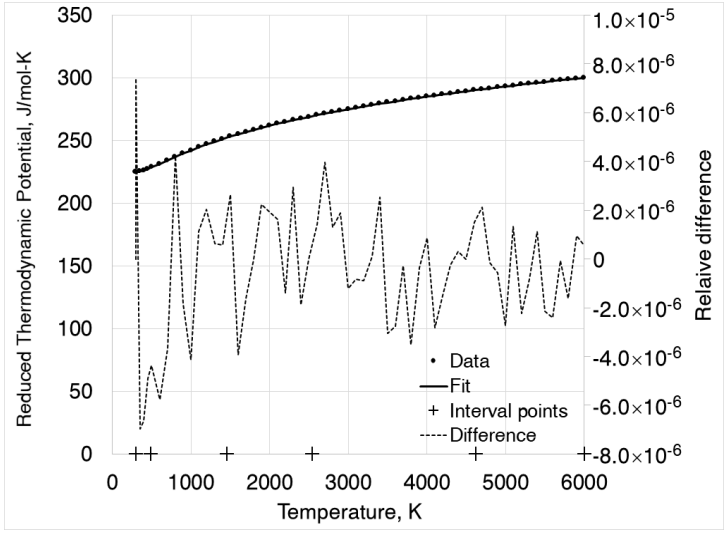
Phosphorus Fluoride (PF) – F1P1(g)

MW = 49.9722 g/mol – $\Delta_f H^o$(298.15 K) = -52.25 kJ/mol – S^o(298.15 K) = 224.968 J/mol-K – p^o = 0.1 MPa



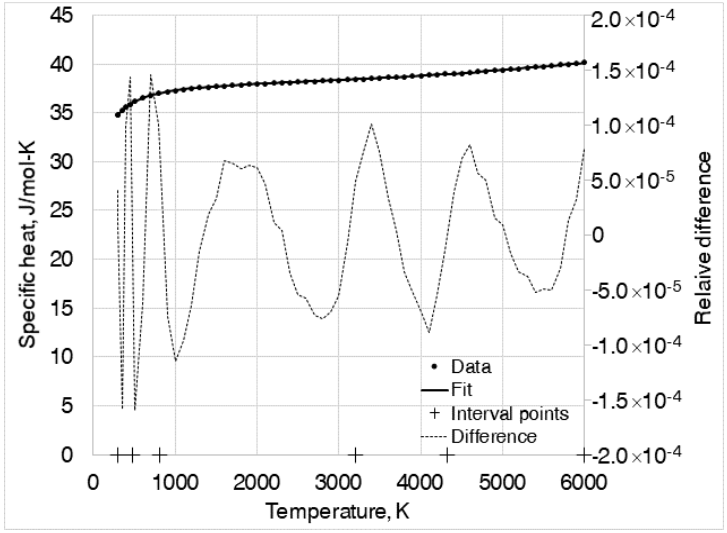


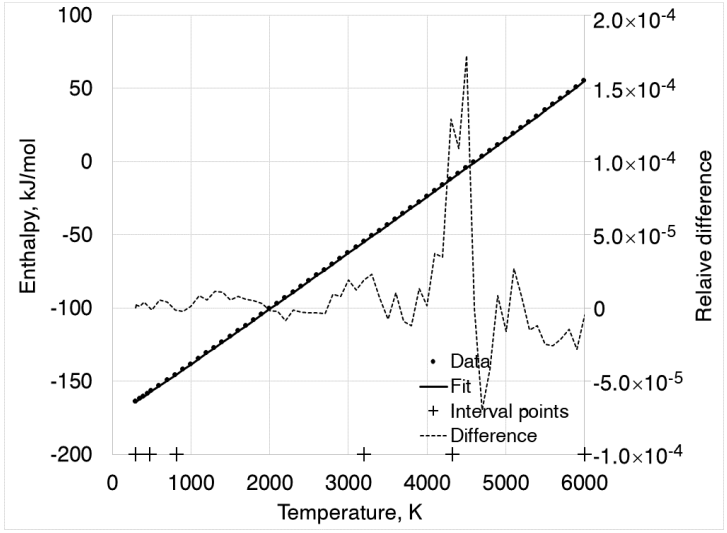


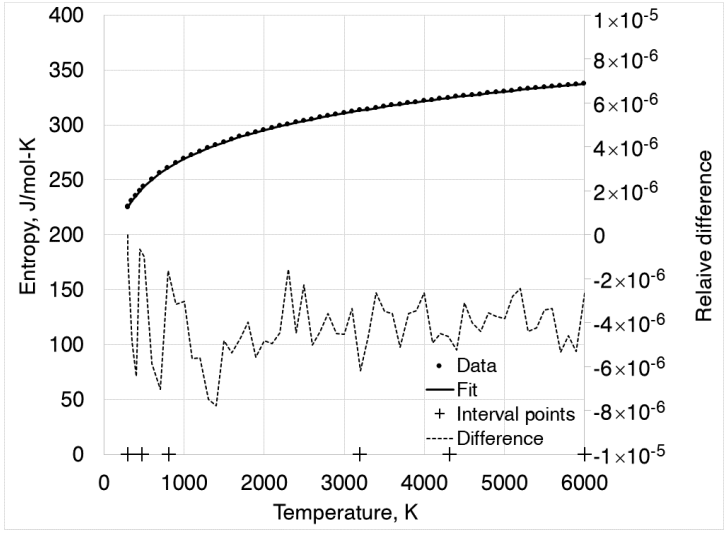


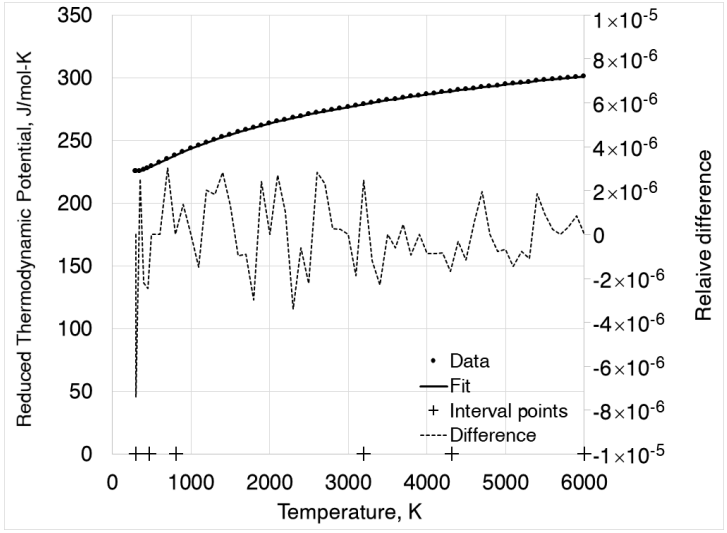
Phosphorus Fluoride, Ion (PF^-) – F1P1-(g)

MW = 49.9727 g/mol – $\Delta_f H^o$(298.15 K) = -164.046 kJ/mol – S^o(298.15 K) = 225.218 J/mol-K – p^o = 0.1 MPa



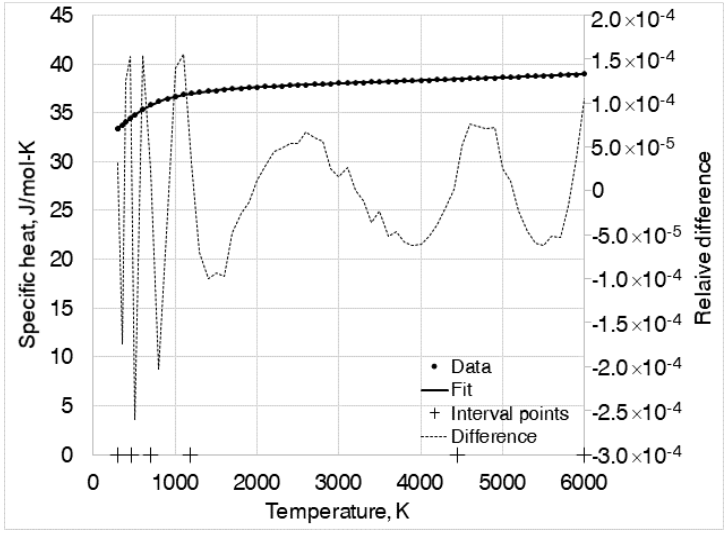


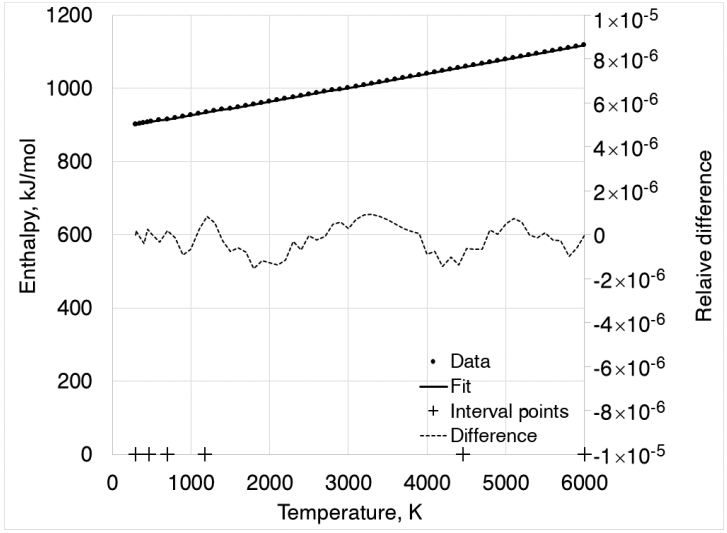


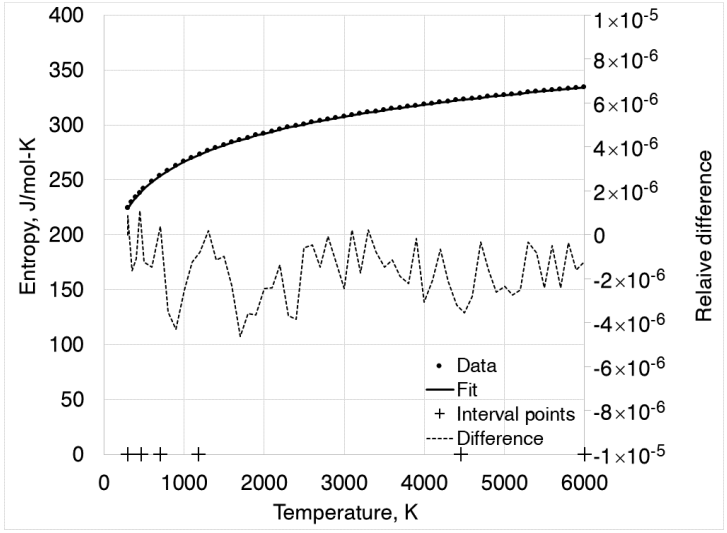


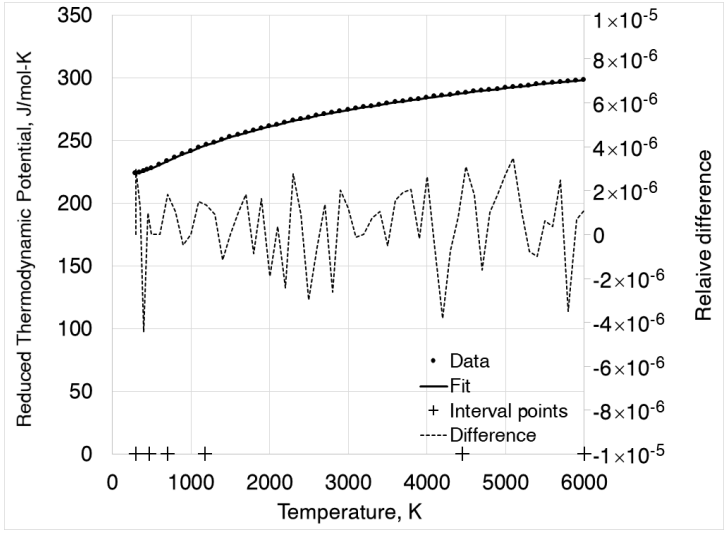
Phosphorus Fluoride, Ion (PF^+) – F1P1+(g)

MW = 49.9716 g/mol – $\Delta_f H^o$(298.15 K) = 901.518 kJ/mol – S^o(298.15 K) = 224.035 J/mol-K – p^o = 0.1 MPa



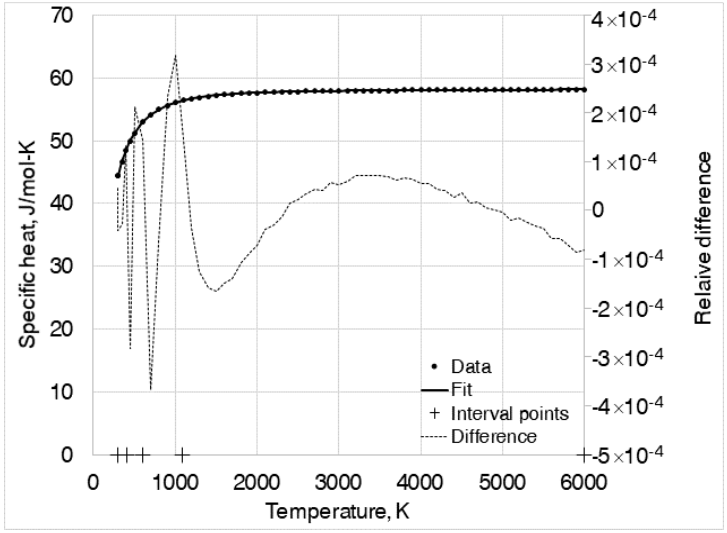


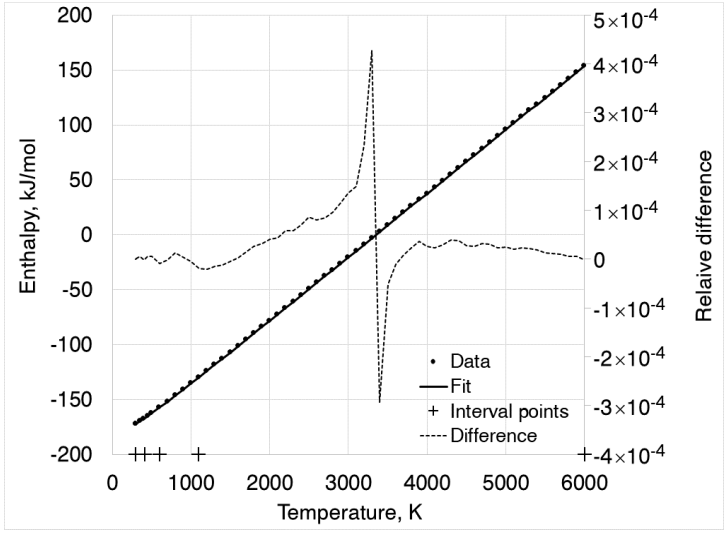


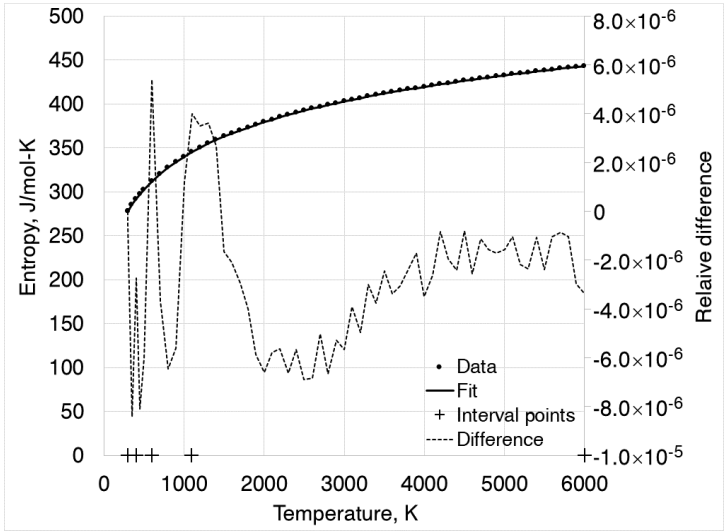


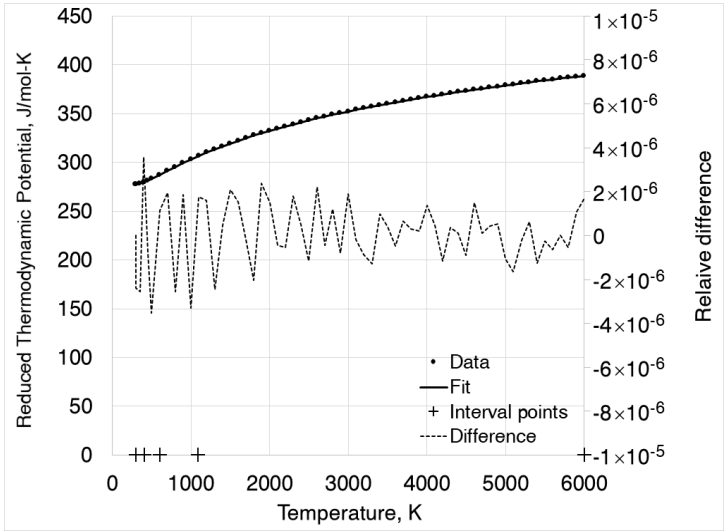
Thiophosphoryl Fluoride (PSF) – F1P1S1(g)

MW = 82.03 g/mol – $\Delta_f H^o$(298.15 K) = -172.268 kJ/mol – S^o(298.15 K) = 277.717 J/mol-K – p^o = 0.1 MPa



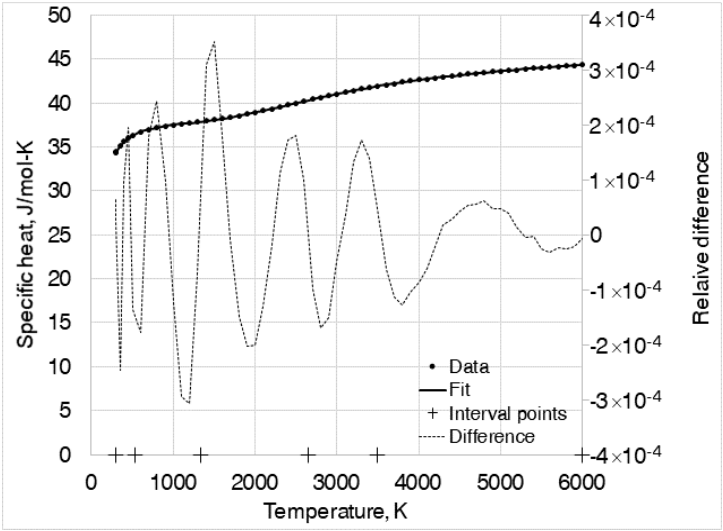


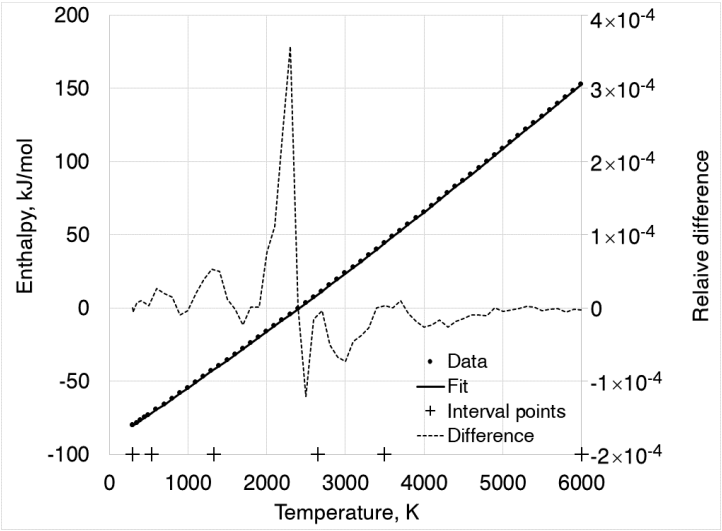


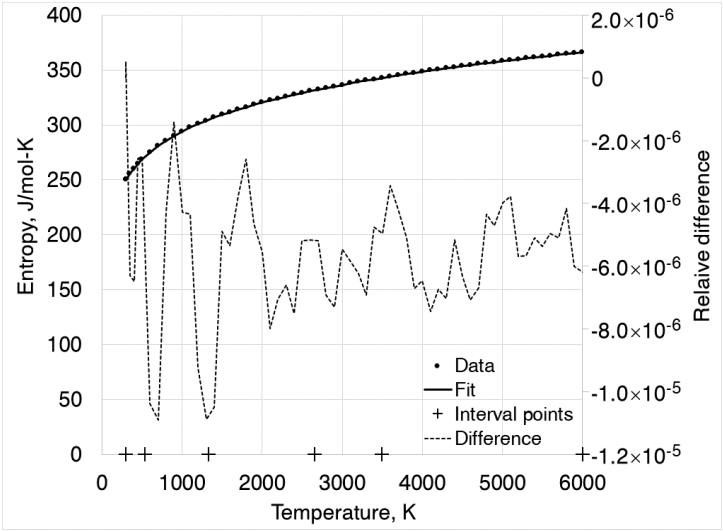


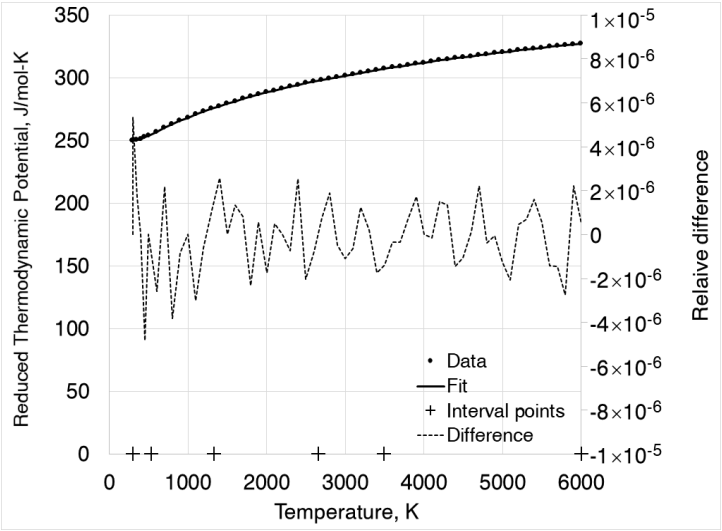
Lead Fluoride (PbF) – F1Pb1(g)

MW = 226.2 g/mol – $\Delta_f H^o$(298.15 K) = -80.266 kJ/mol – S^o(298.15 K) = 249.957 J/mol-K – p^o = 0.1 MPa



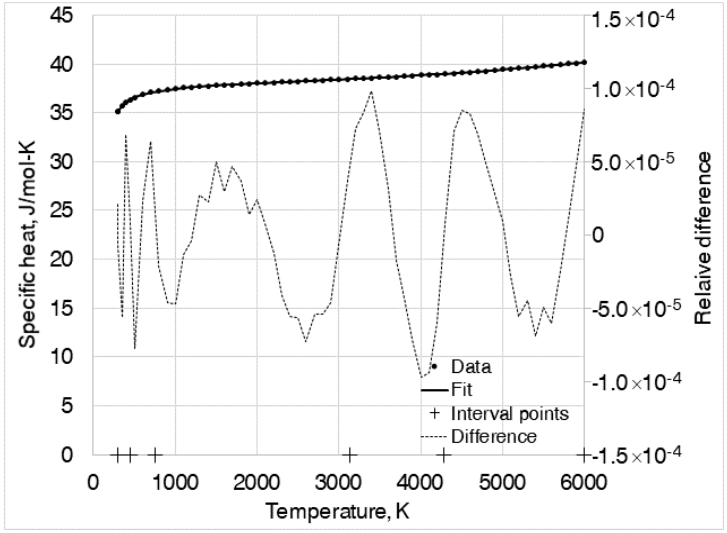


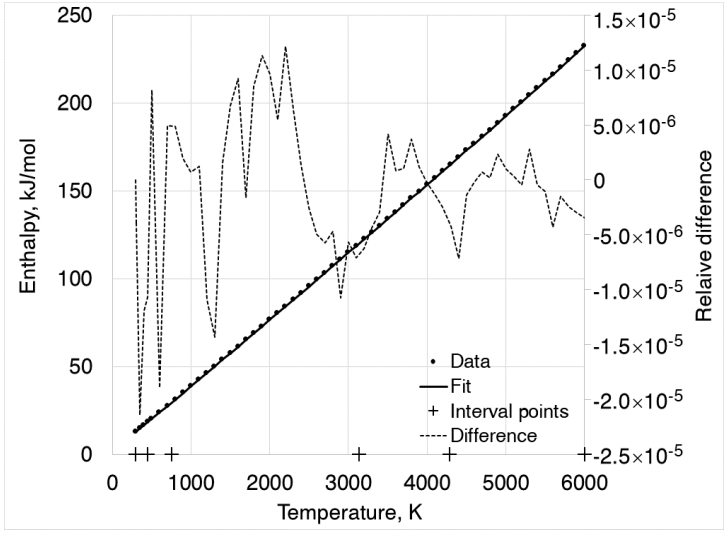


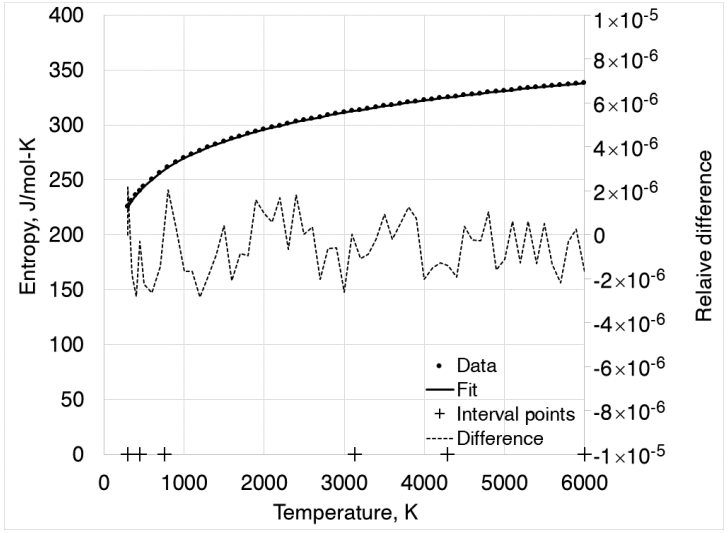


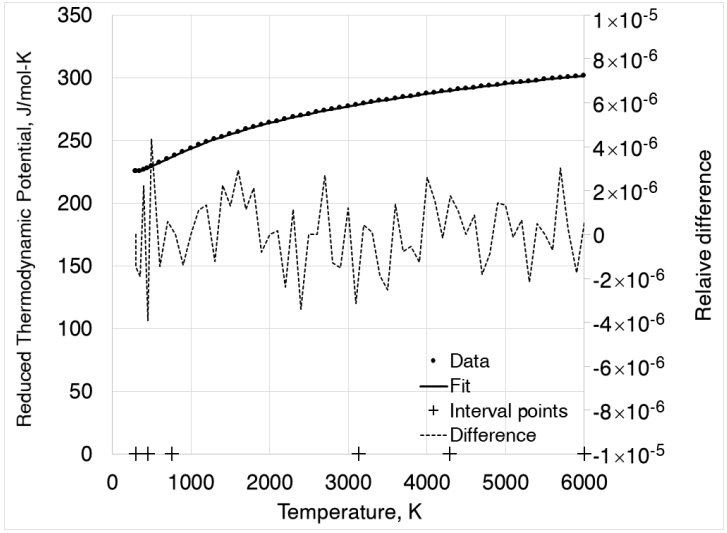
Sulfur Fluoride (SF) – F1S1(g)

MW = 51.06 g/mol – $\Delta_f H^o$(298.15 K) = 12.97 kJ/mol – S^o(298.15 K) = 225.275 J/mol-K – p^o = 0.1 MPa



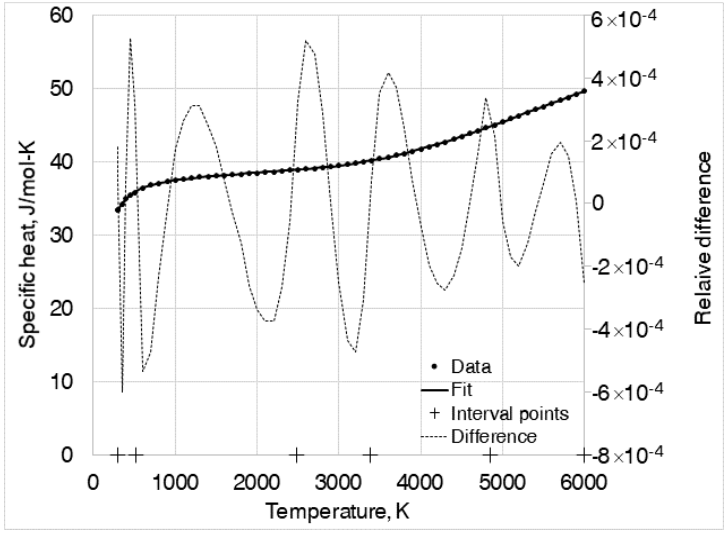


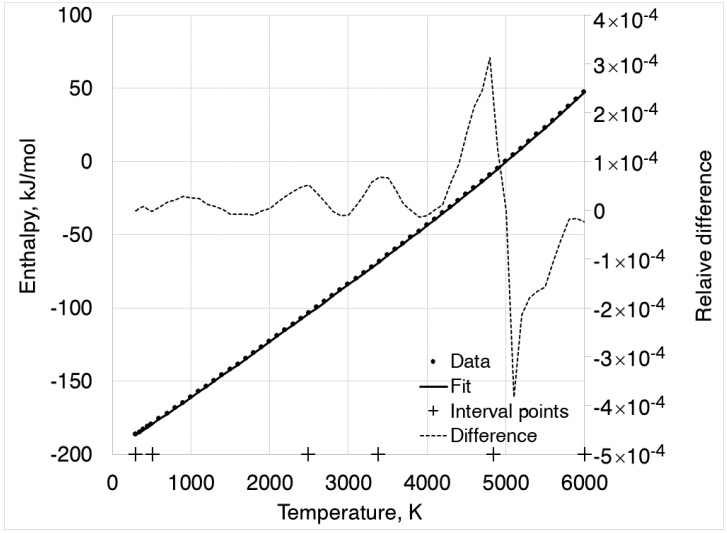


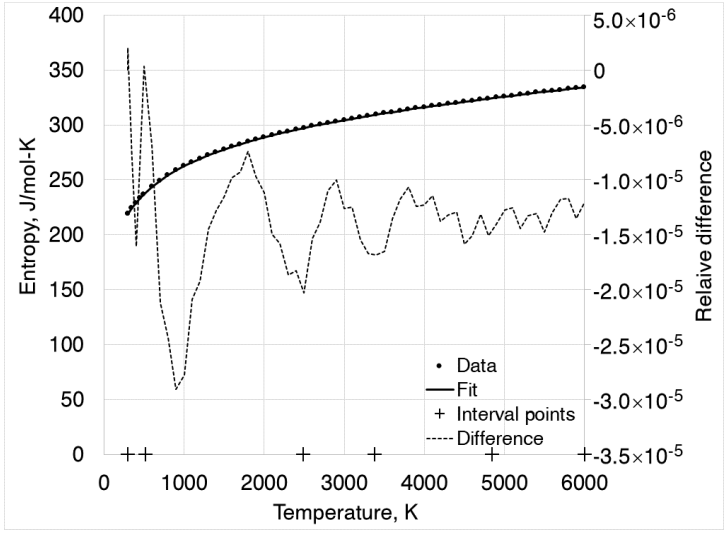


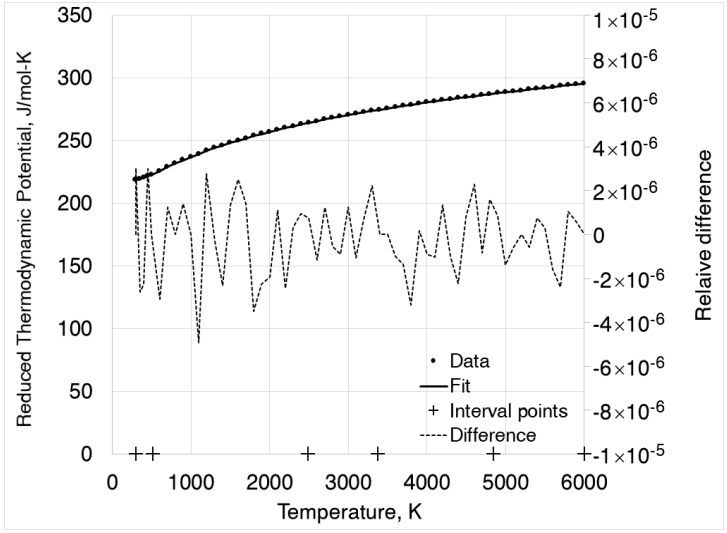
Sulfur Fluoride, Ion (SF^-) – F1S1-(g)

MW = 51.06 g/mol – $\Delta_f H^o$(298.15 K) = -186.59 kJ/mol – S^o(298.15 K) = 218.956 J/mol-K – p^o = 0.1 MPa



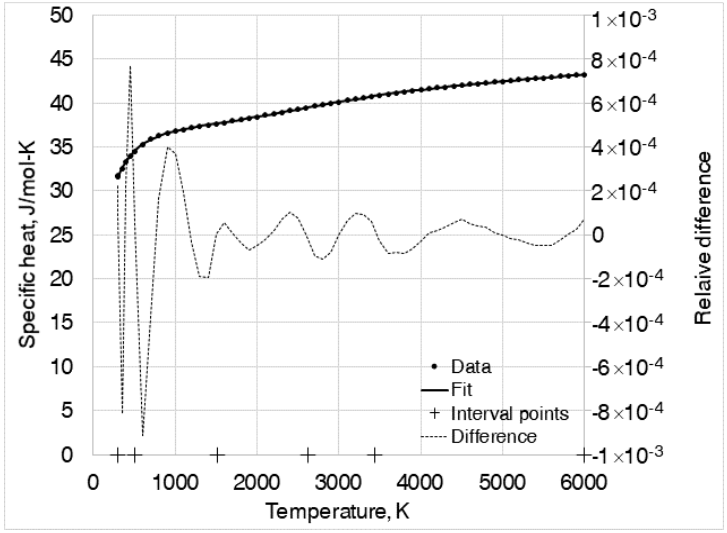


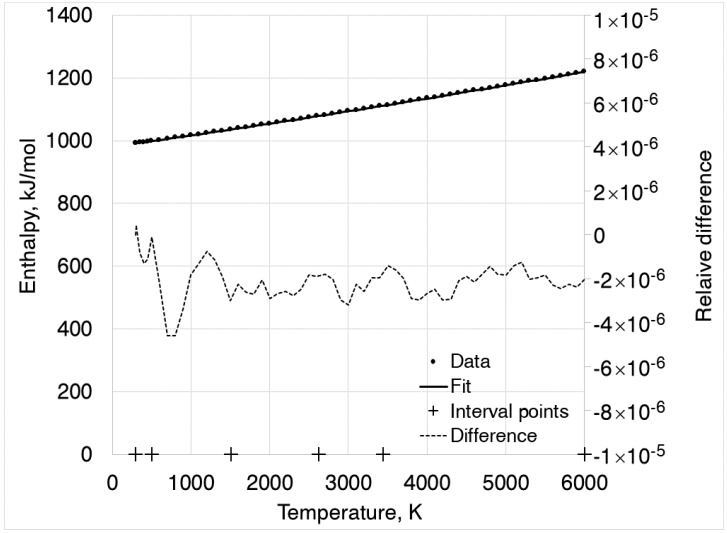


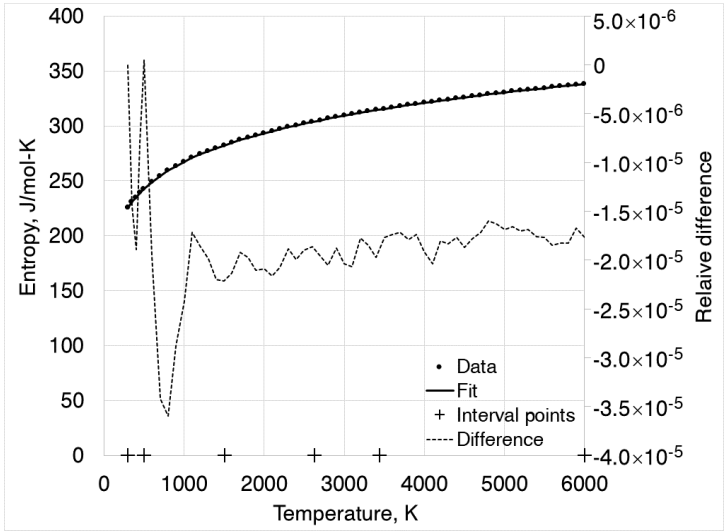


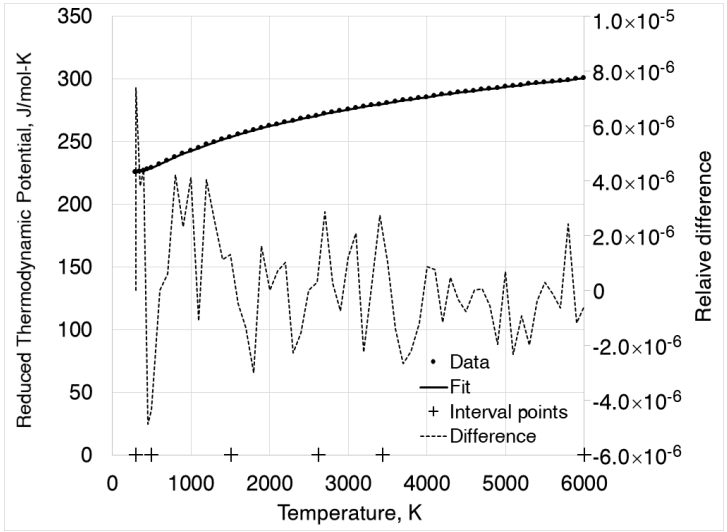
Sulfur Fluoride, Ion (SF^+) – F1S1+(g)

MW = 51.06 g/mol – $\Delta_f H^o$(298.15 K) = 991.901 kJ/mol – S^o(298.15 K) = 225.405 J/mol-K – p^o = 0.1 MPa



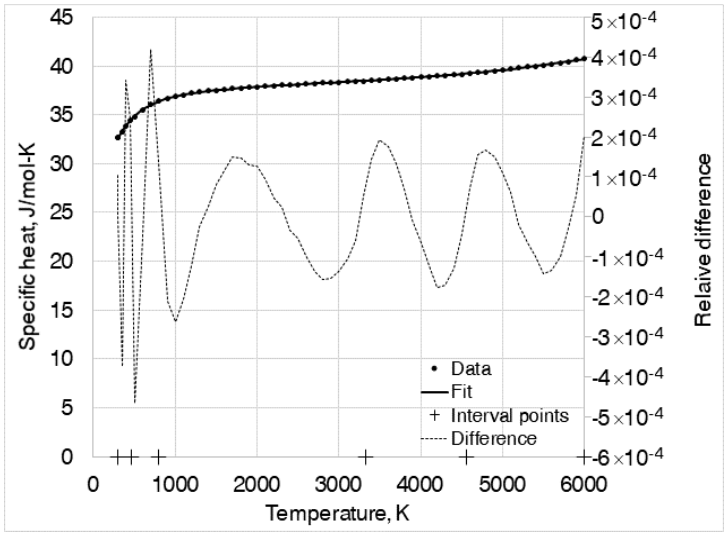


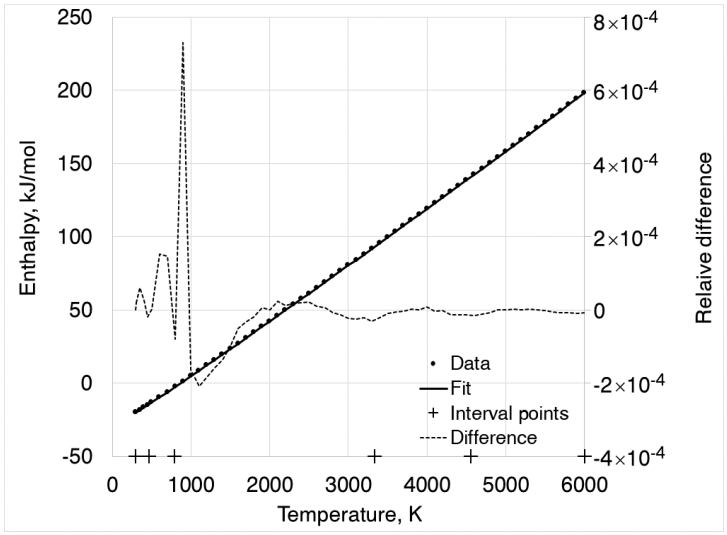


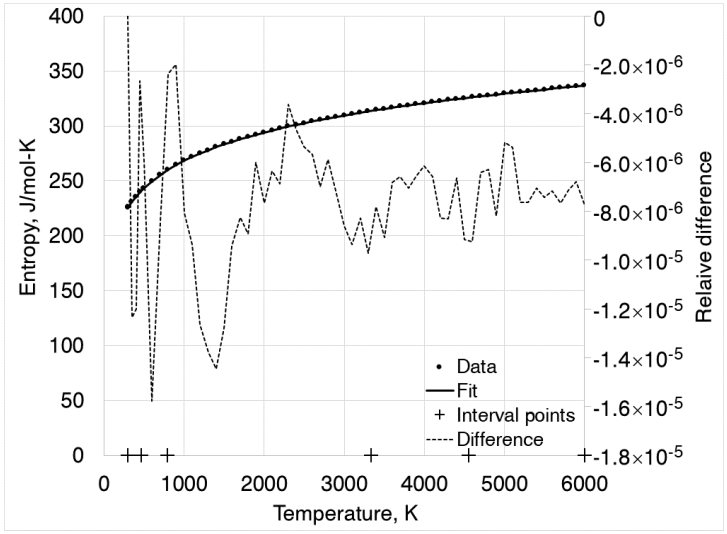


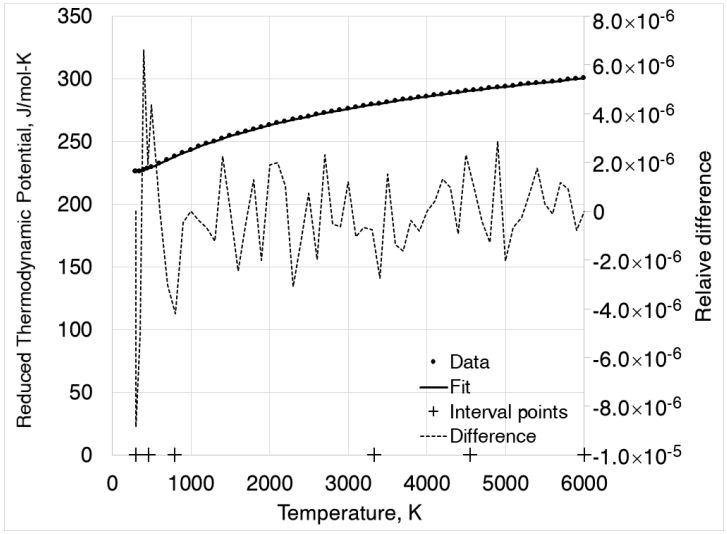
Fluorosilylidyne (SiF) – F1Si1(g)

MW = 47.0839 g/mol – $\Delta_f H^o$(298.15 K) = -20.083 kJ/mol – S^o(298.15 K) = 225.786 J/mol-K – p^o = 0.1 MPa



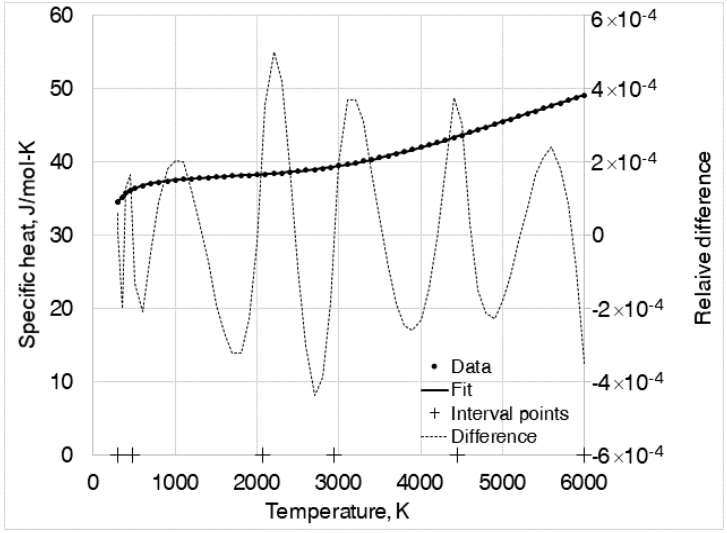


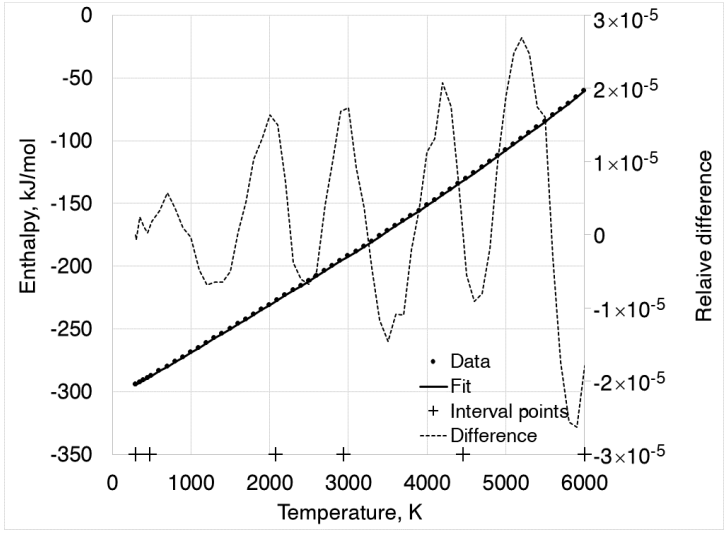


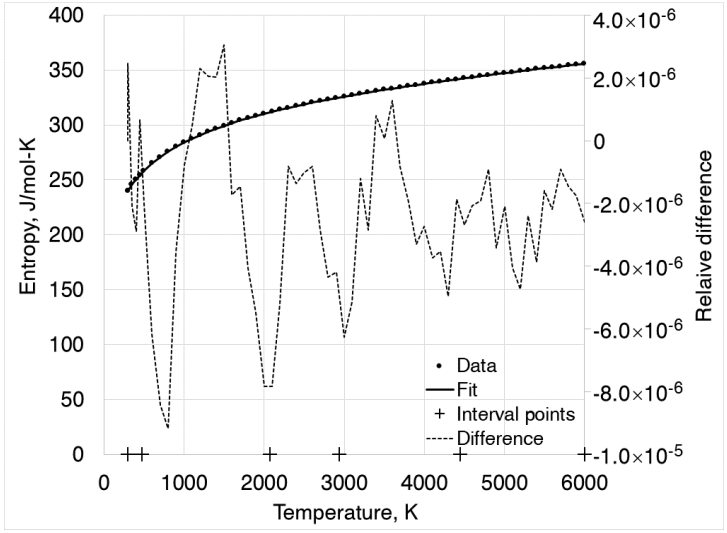


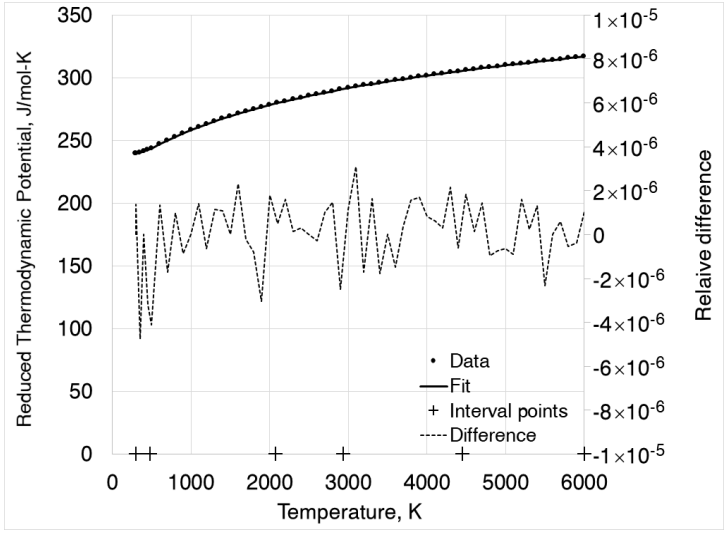
Strontium Fluoride (SrF) – F1Sr1(g)

MW = 106.62 g/mol – $\Delta_f H^o$(298.15 K) = -294.554 kJ/mol – S^o(298.15 K) = 239.919 J/mol-K – p^o = 0.1 MPa



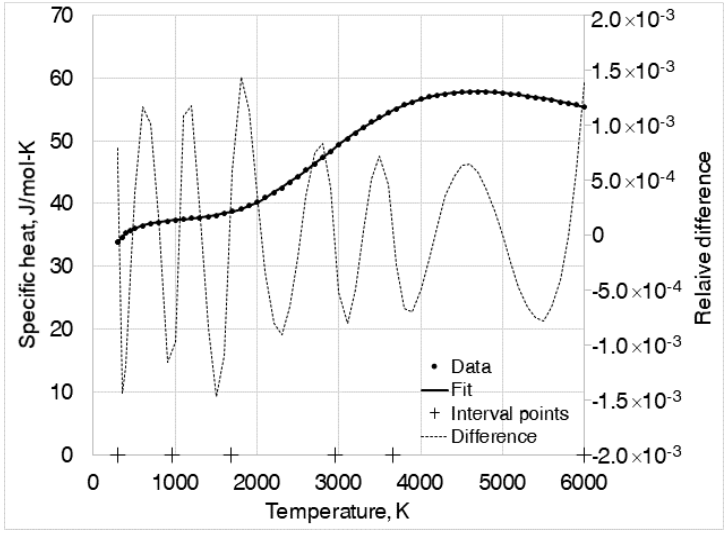


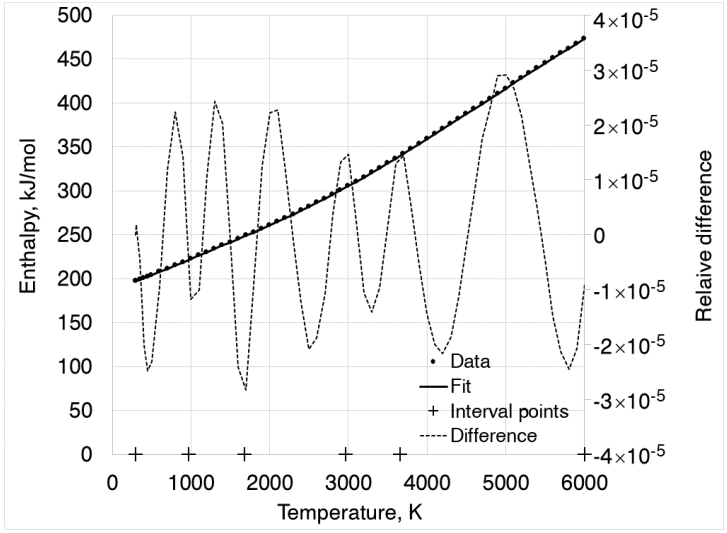


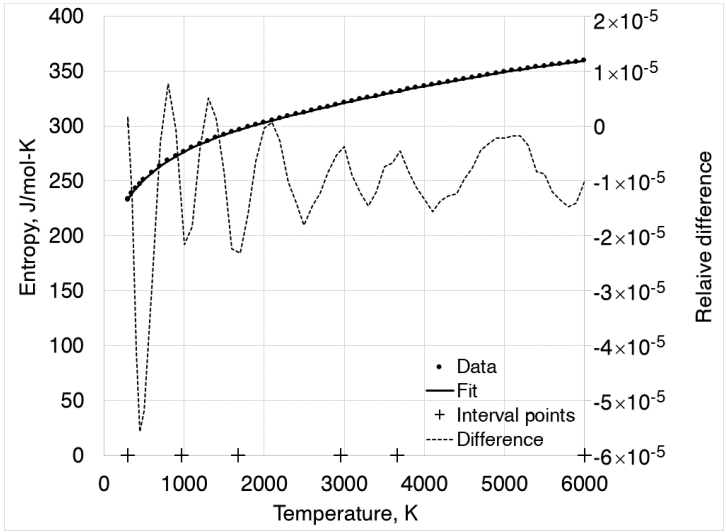


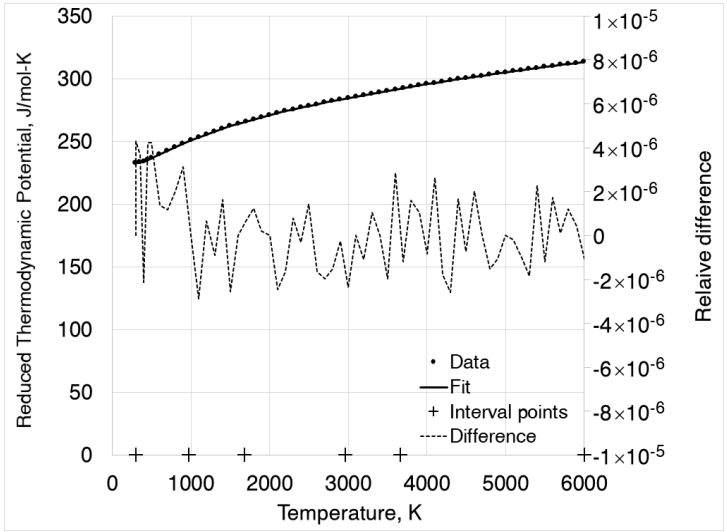
Strontium Fluoride, Ion (SrF^+) – F1Sr1+(g)

MW = 106.62 g/mol – $\Delta_f H^o$(298.15 K) = 197.066 kJ/mol – S^o(298.15 K) = 232.846 J/mol-K – p^o = 0.1 MPa



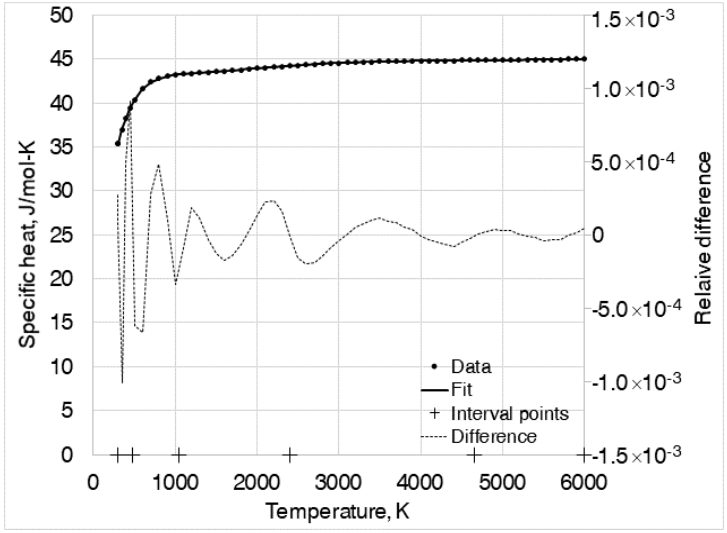


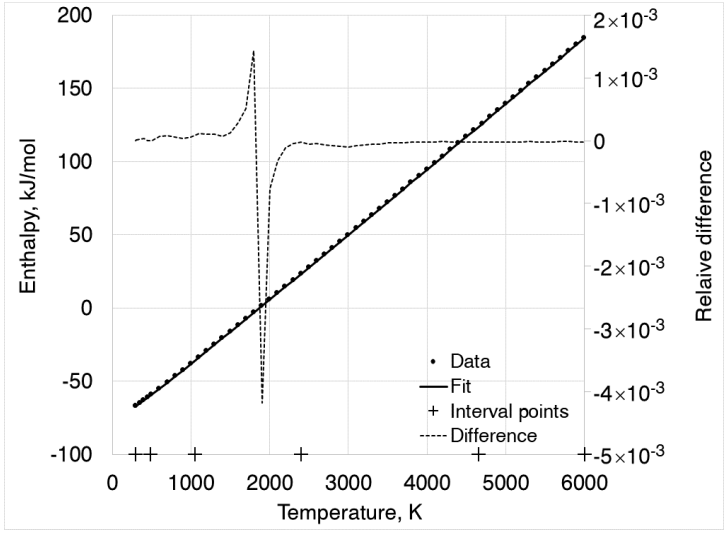


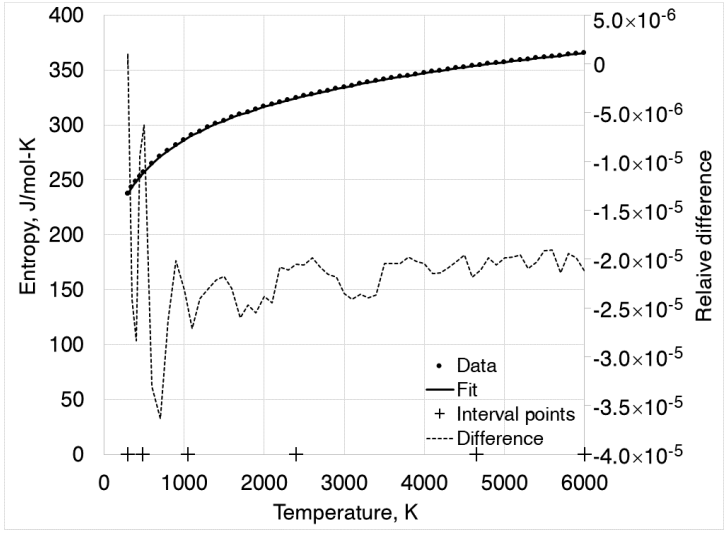


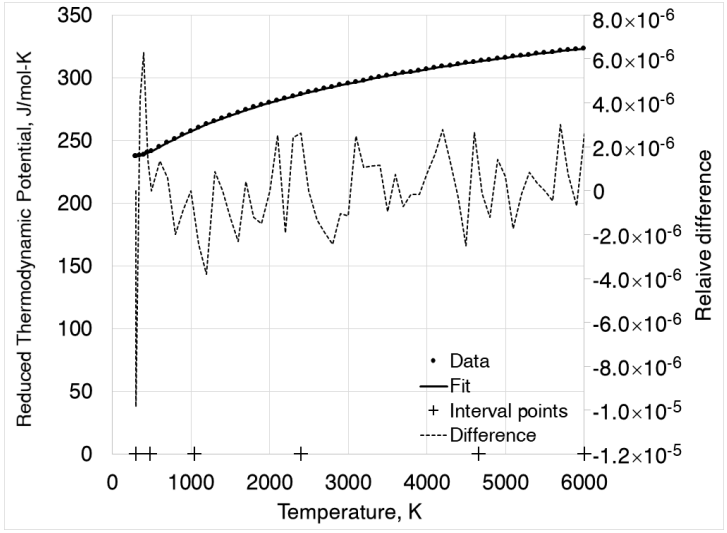
Titanium Fluoride (TiF) – F1Ti1(g)

MW = 66.88 g/mol – $\Delta_f H^o$(298.15 K) = -66.944 kJ/mol – S^o(298.15 K) = 237.309 J/mol-K – p^o = 0.1 MPa



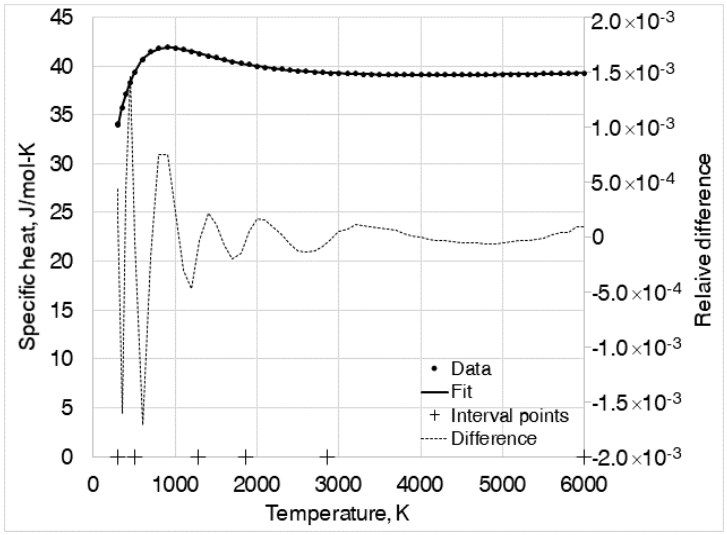


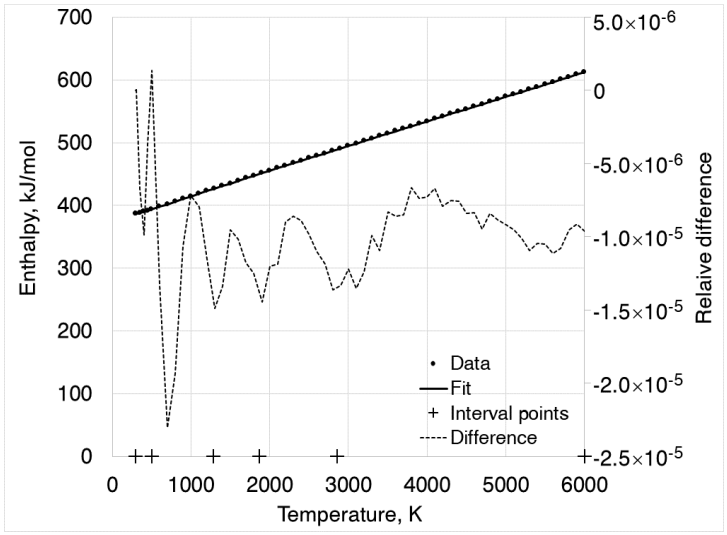


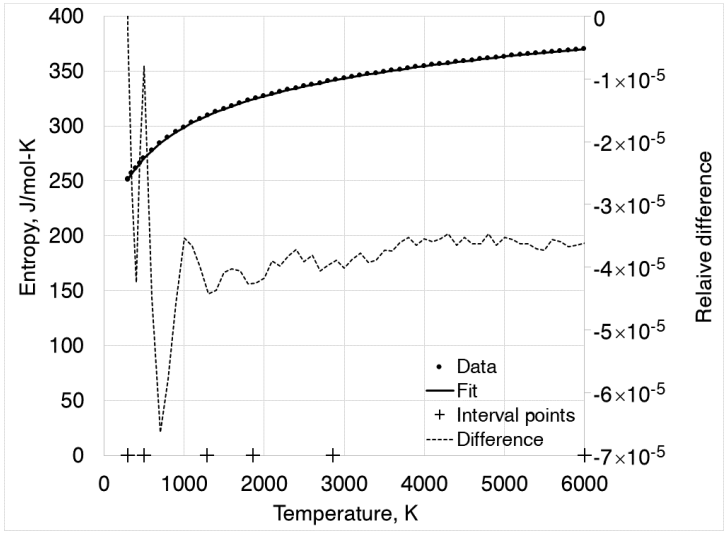


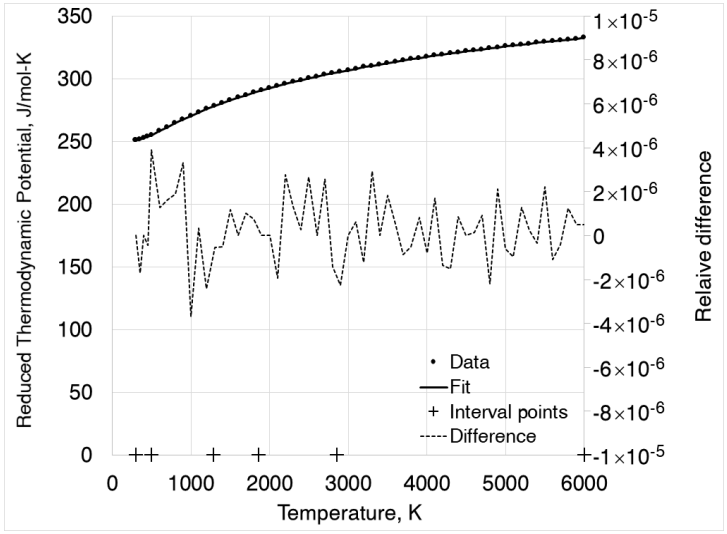
Tungsten Fluoride (WF) – F1W1(g)

MW = 202.85 g/mol – $\Delta_f H^o$(298.15 K) = 386.183 kJ/mol – S^o(298.15 K) = 251.141 J/mol-K – p^o = 0.1 MPa



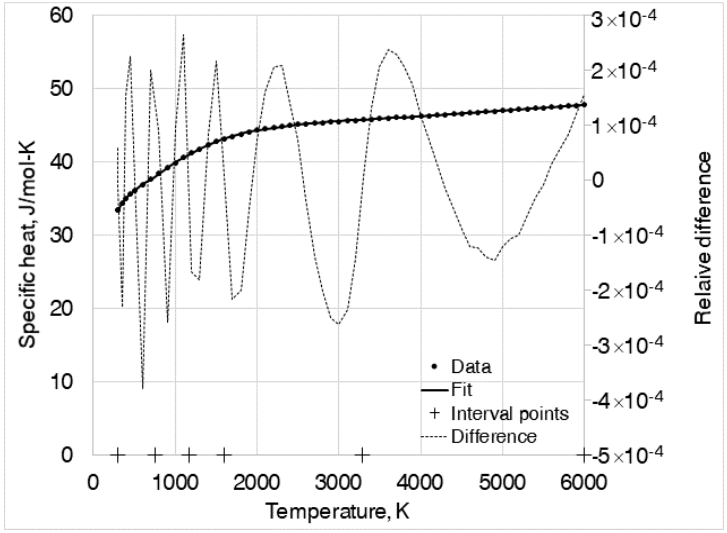


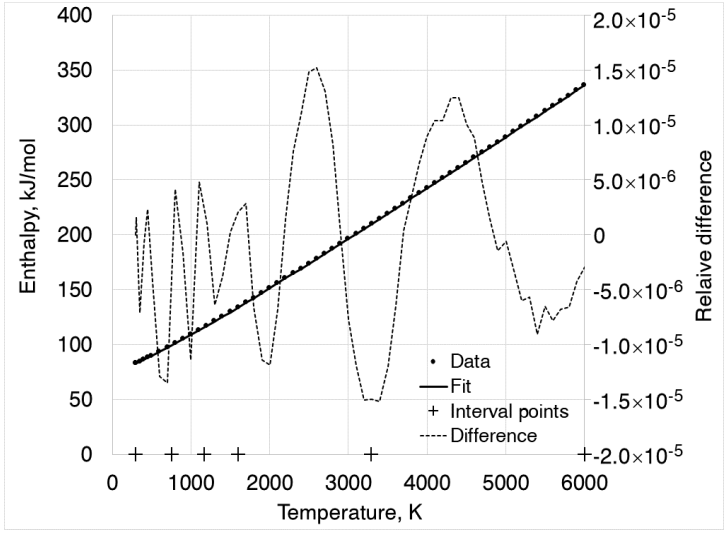


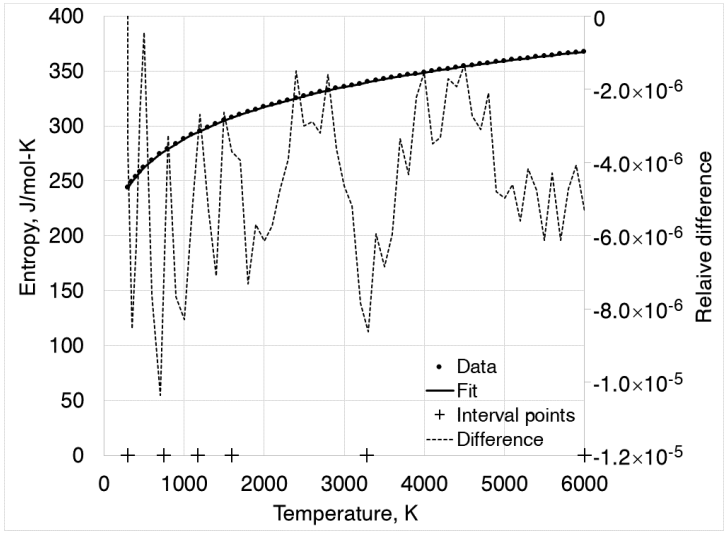


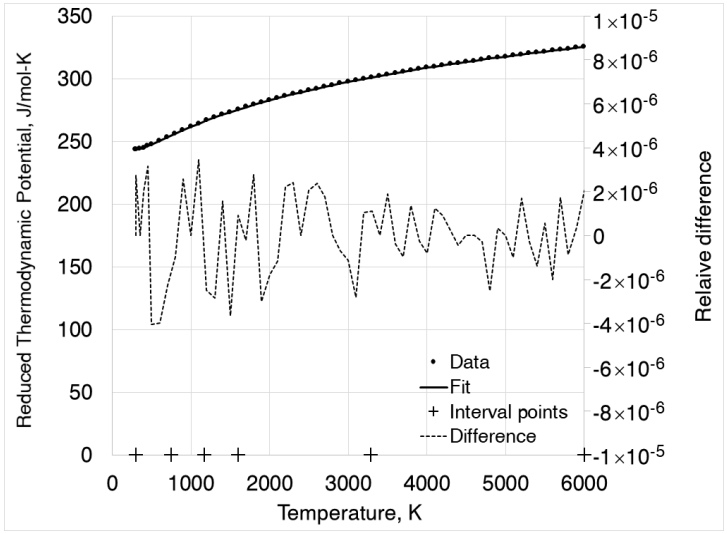
Zirconium Fluoride (ZrF) – F1Zr1(g)

MW = 110.22 g/mol – $\Delta_f H^o$(298.15 K) = 82.843 kJ/mol – S^o(298.15 K) = 243.736 J/mol-K – p^o = 0.1 MPa



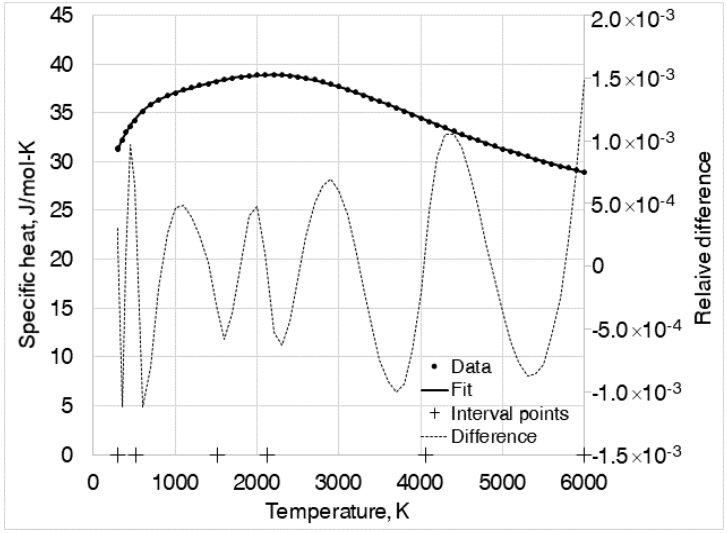


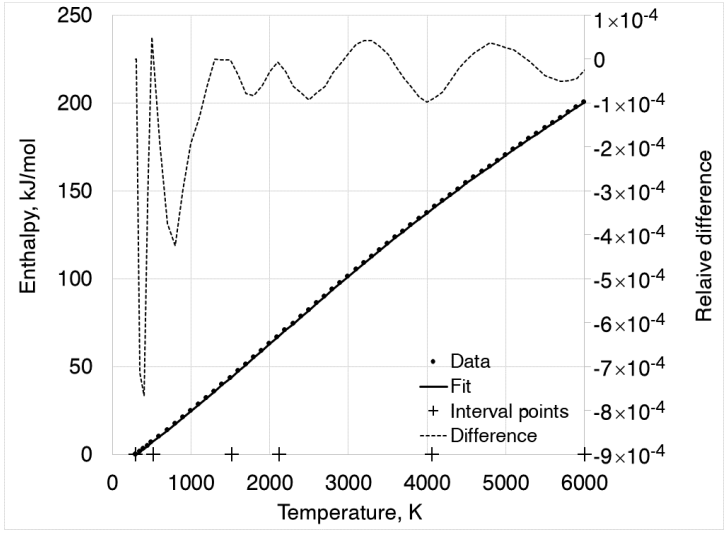


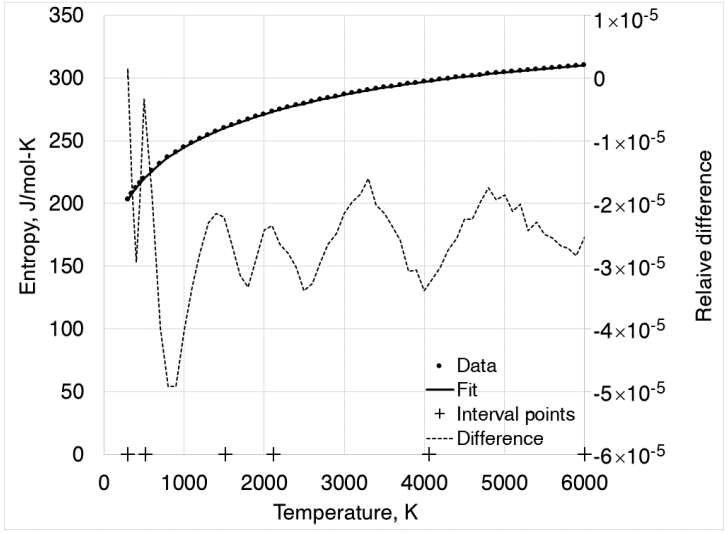


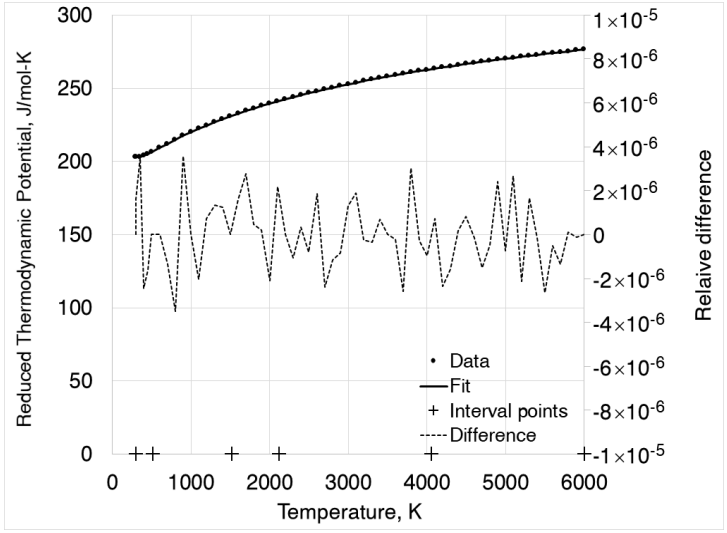
Fluorine (F_2) – F2(ref)

MW = 37.9968 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 202.789 J/mol-K – p^o = 0.1 MPa



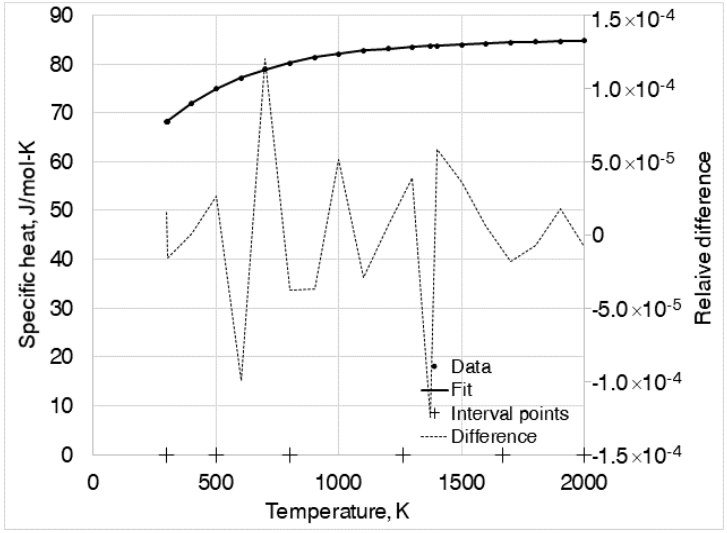


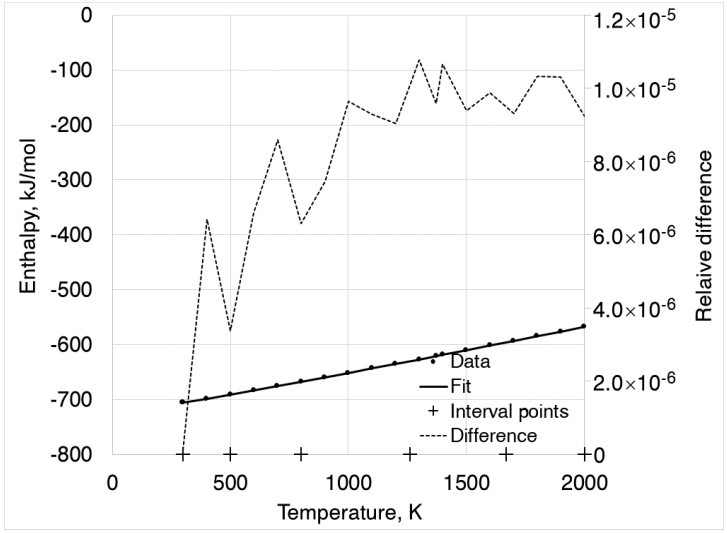


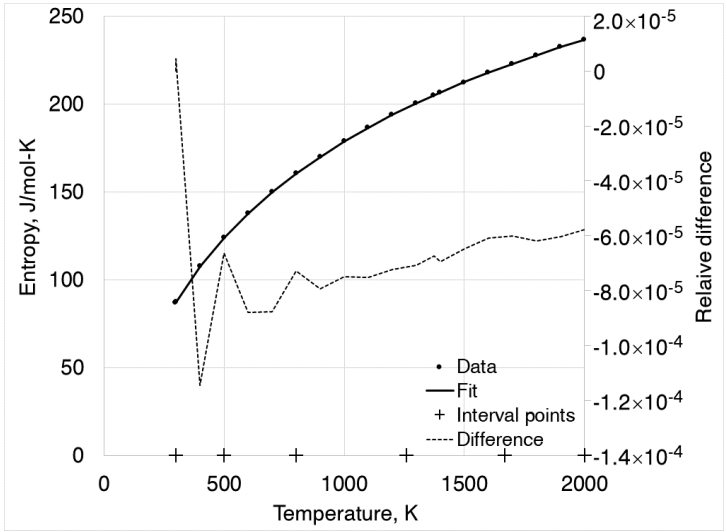


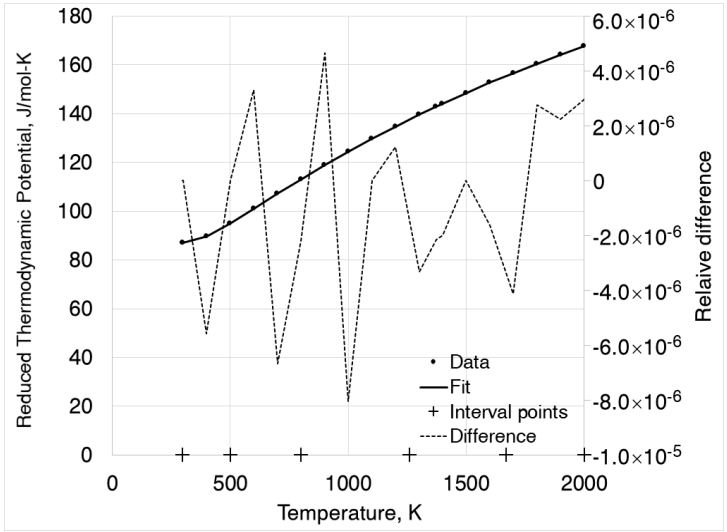
Iron Fluoride (FeF_2) – F2Fe1(cr)

MW = 93.844 g/mol – $\Delta_f H^o$(298.15 K) = -705.841 kJ/mol – S^o(298.15 K) = 86.985 J/mol-K – p^o = 0.1 MPa



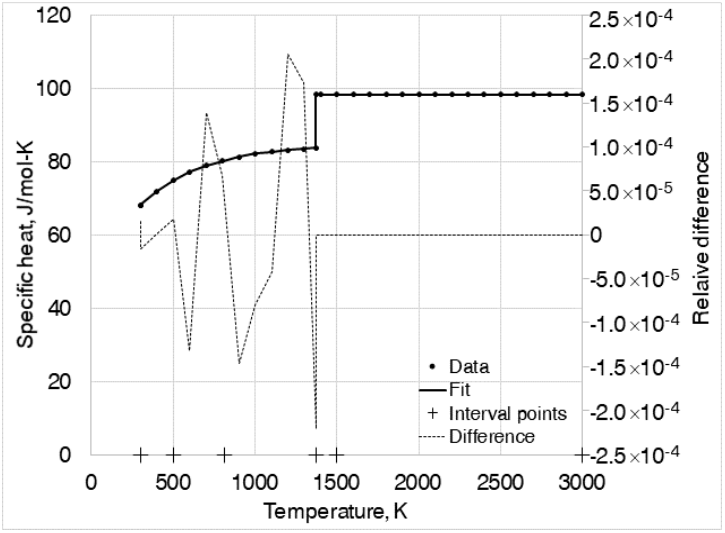


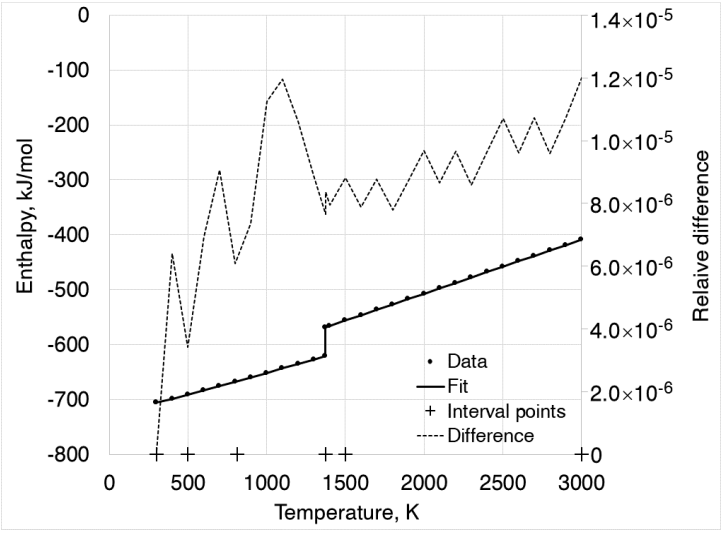


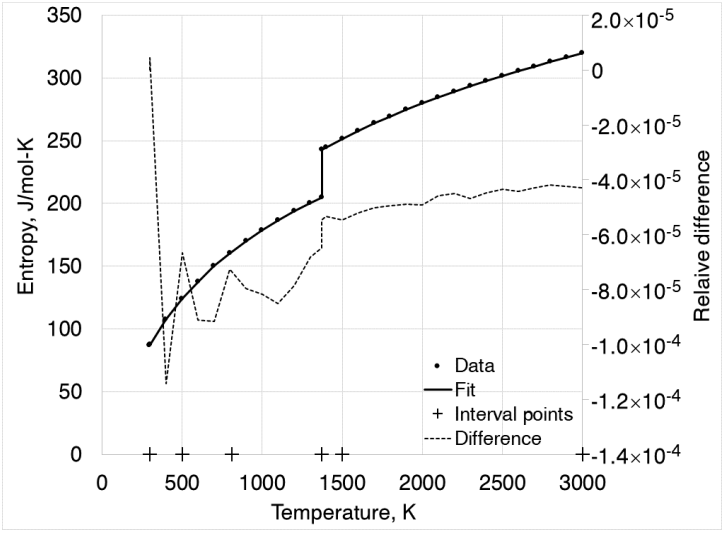


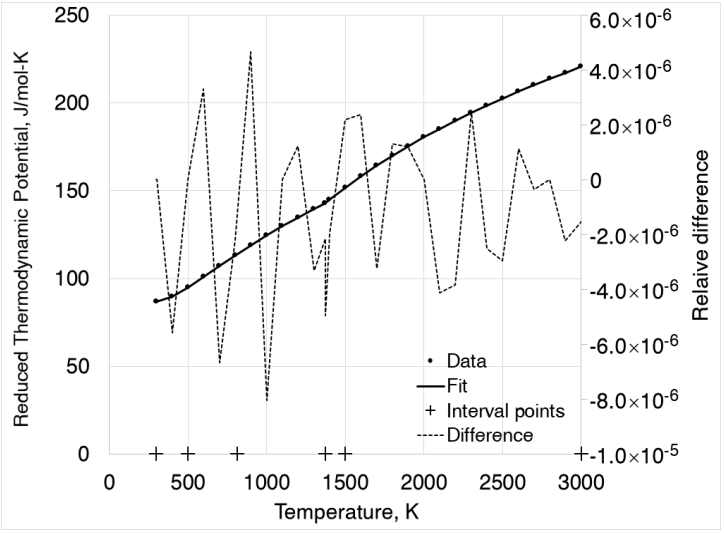
Iron Fluoride (FeF_2) – F2Fe1(cr,l)

MW = 93.844 g/mol – $\Delta_f H^o$(298.15 K) = -705.841 kJ/mol – S^o(298.15 K) = 86.985 J/mol-K – p^o = 0.1 MPa



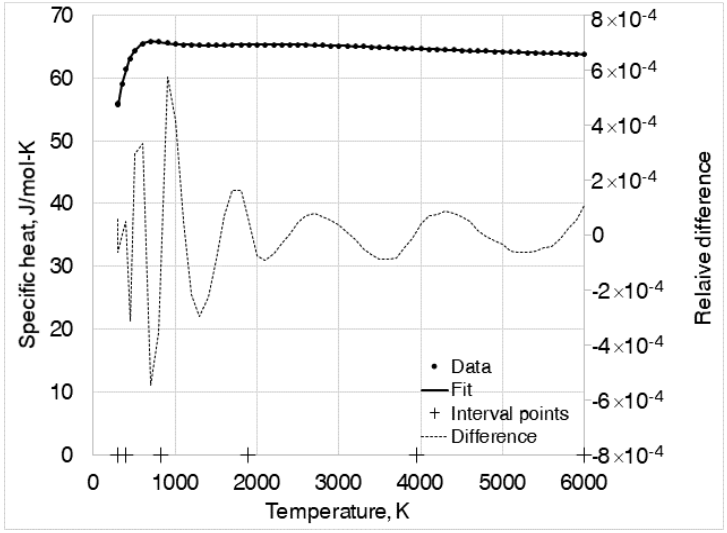


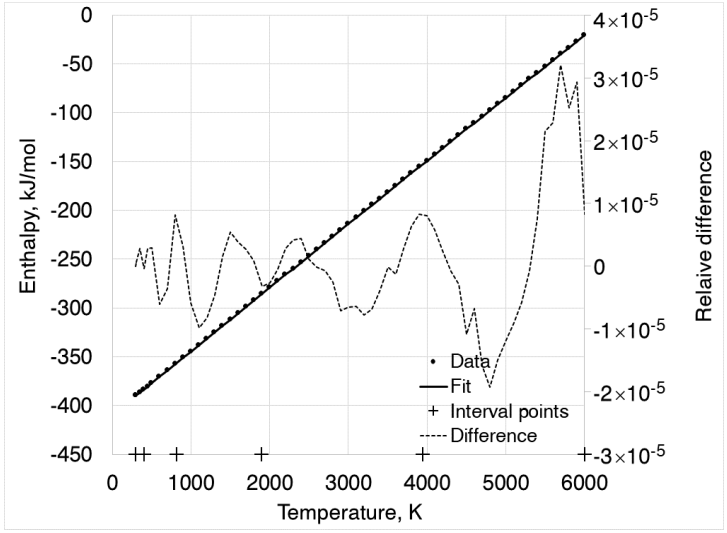


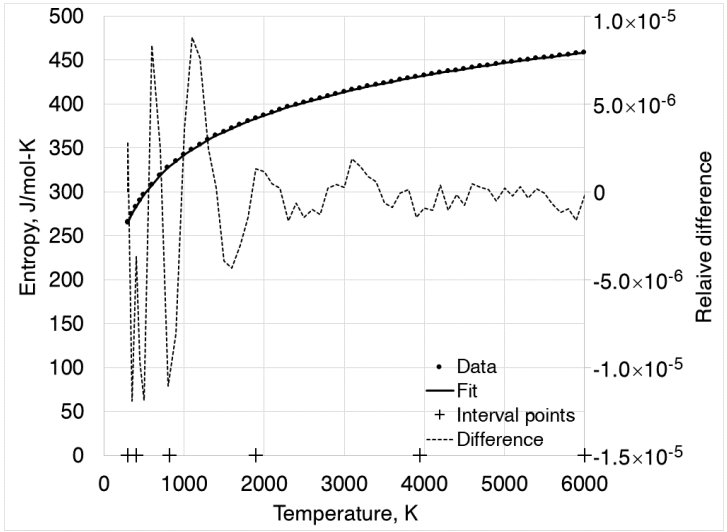


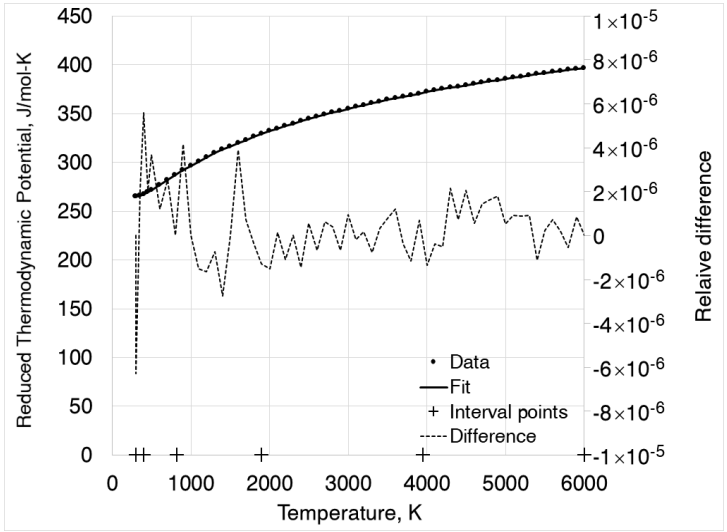
Iron Fluoride (FeF_2) – F2Fe1(g)

MW = 93.844 g/mol – $\Delta_f H^o$(298.15 K) = -389.53 kJ/mol – S^o(298.15 K) = 265.205 J/mol-K – p^o = 0.1 MPa



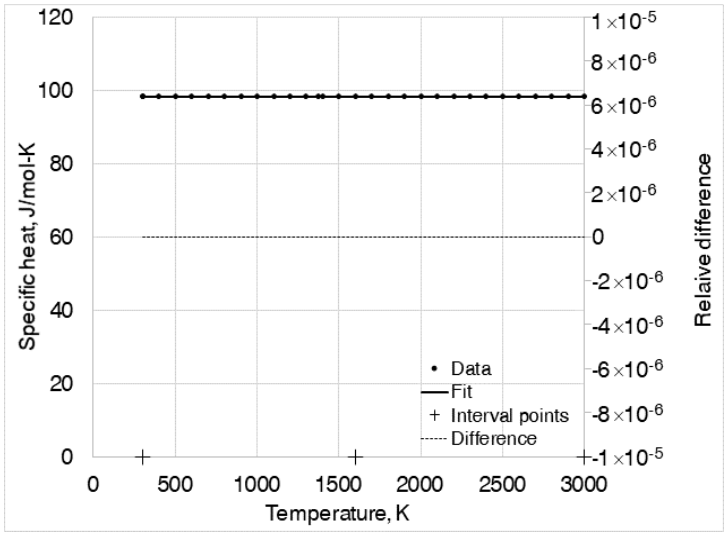


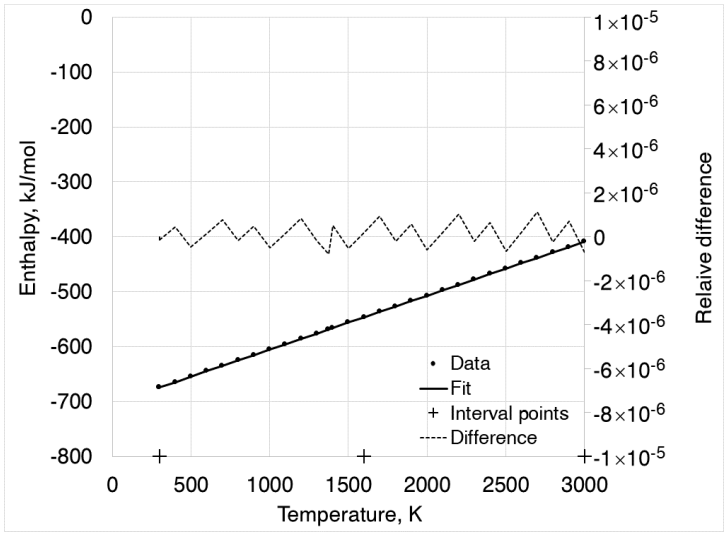


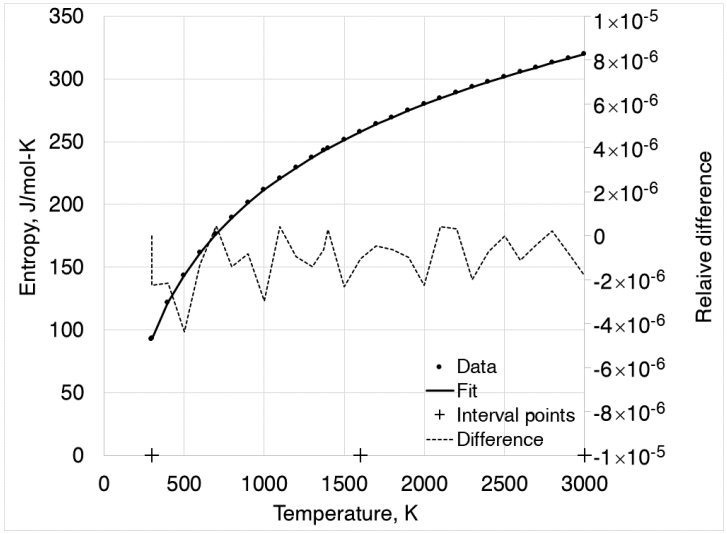


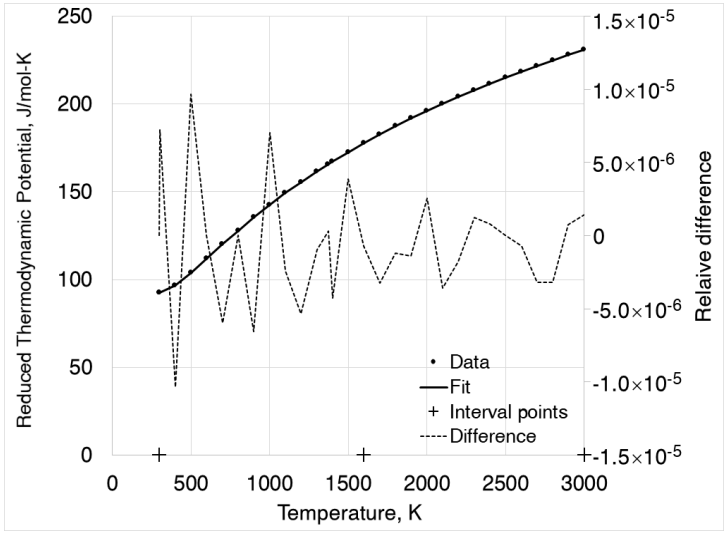
Iron Fluoride (FeF_2) – F2Fe1(l)

MW = 93.844 g/mol – $\Delta_f H^o$(298.15 K) = -674.549 kJ/mol – S^o(298.15 K) = 92.537 J/mol-K – p^o = 0.1 MPa



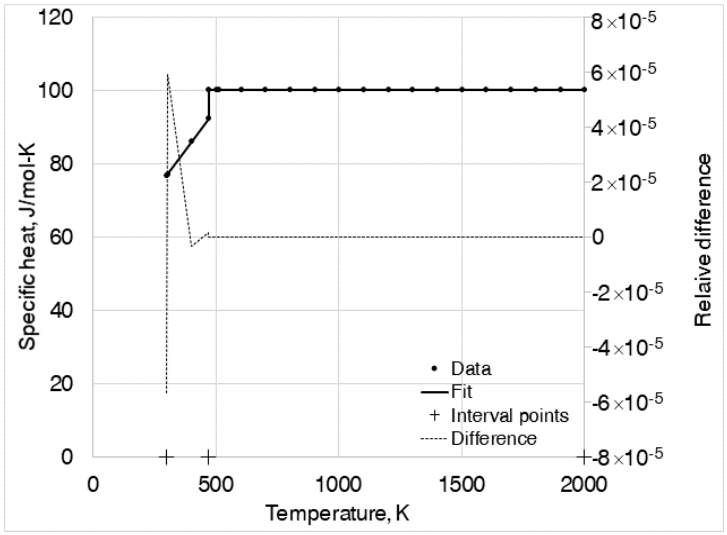


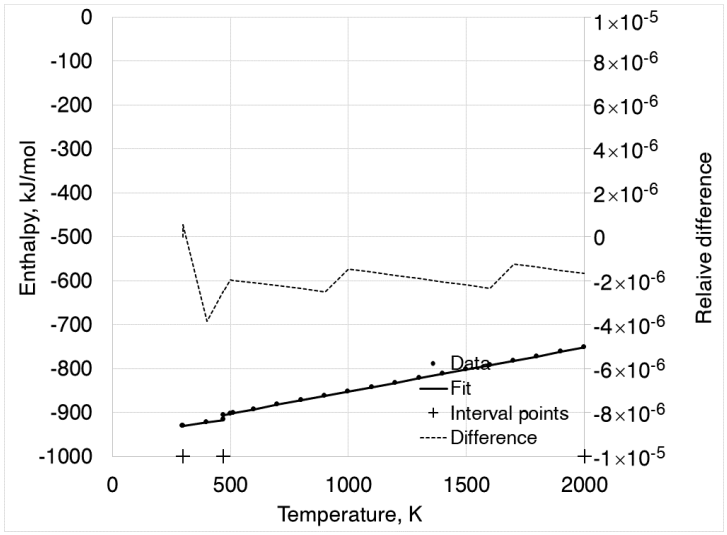


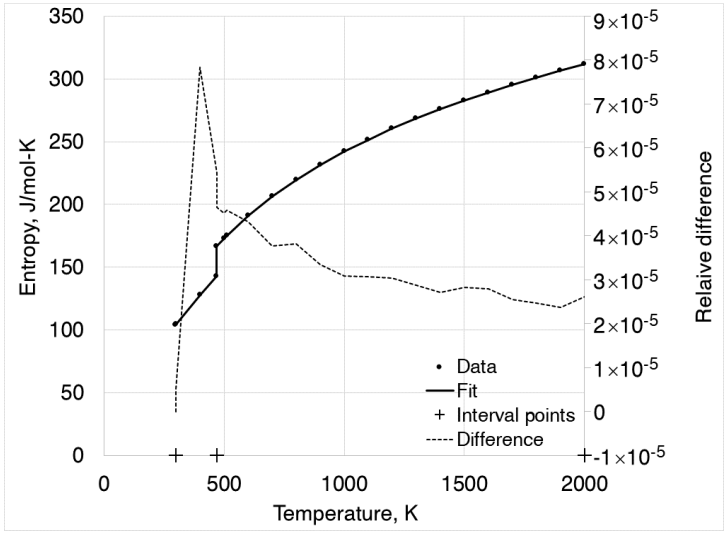


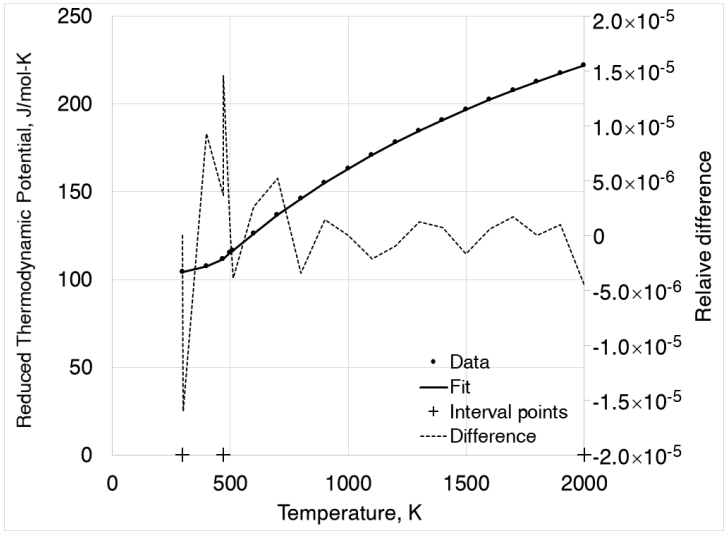
Potassium Fluoride ($K(HF_2)$) – F2H1K1(cr)

MW = 78.103 g/mol – $\Delta_f H^o$(298.15 K) = -931.233 kJ/mol – S^o(298.15 K) = 104.264 J/mol-K – p^o = 0.1 MPa



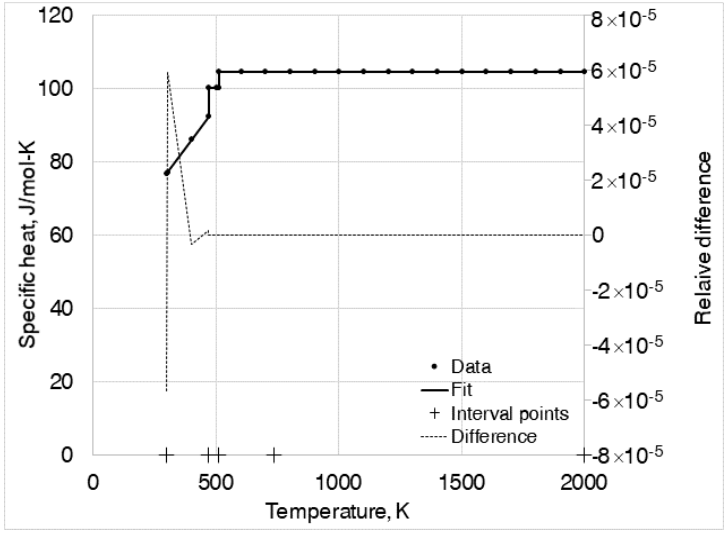


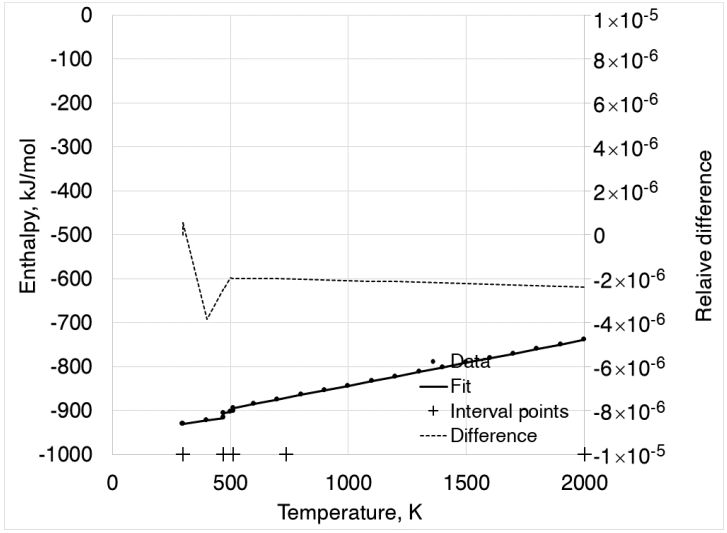


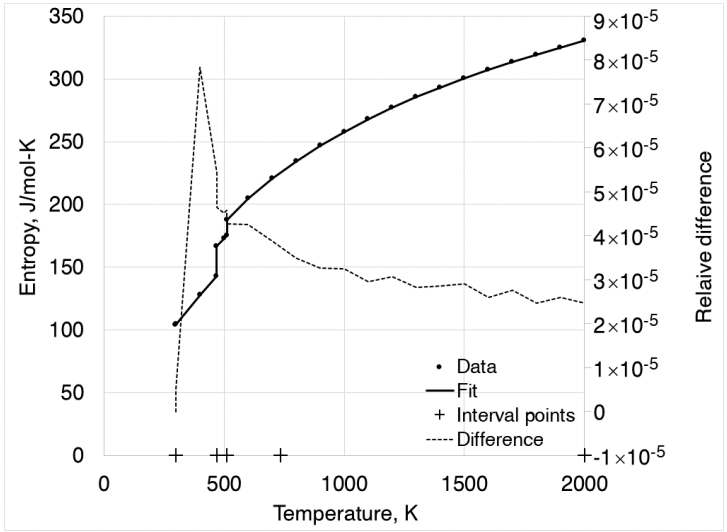


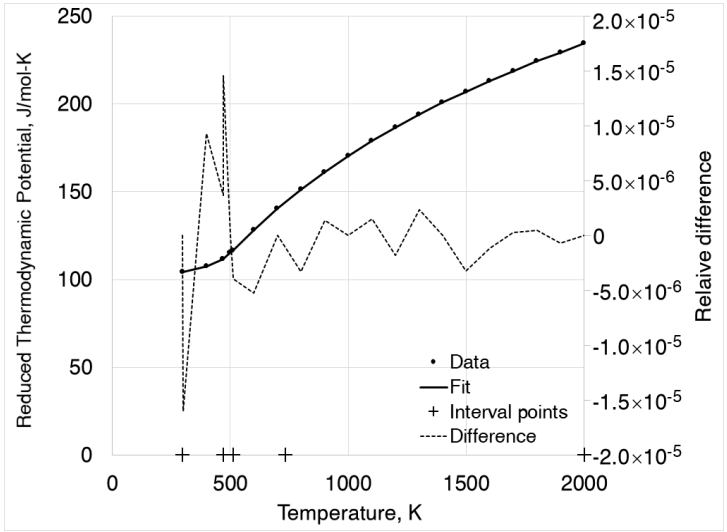
Potassium Fluoride ($K(HF_2)$) – F2H1K1(cr,l)

MW = 78.103 g/mol – $\Delta_f H^o$(298.15 K) = -931.233 kJ/mol – S^o(298.15 K) = 104.264 J/mol-K – p^o = 0.1 MPa



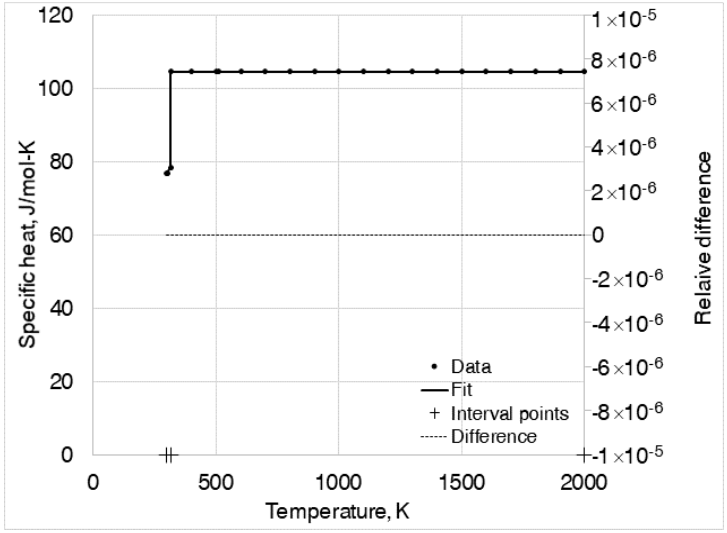


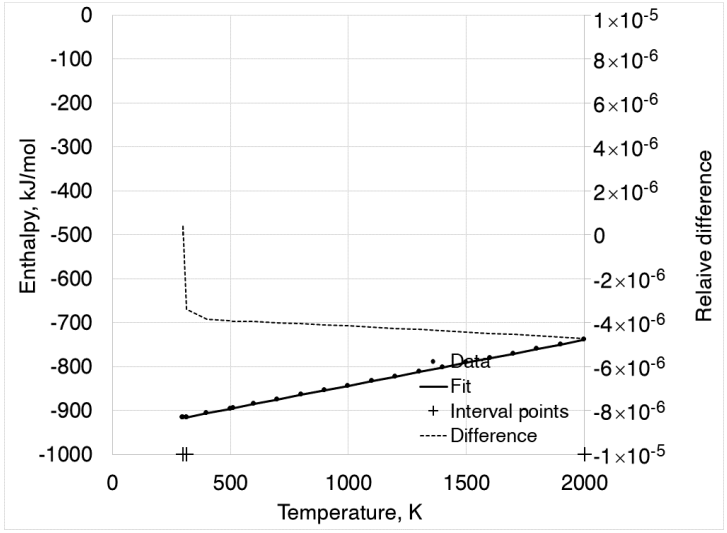


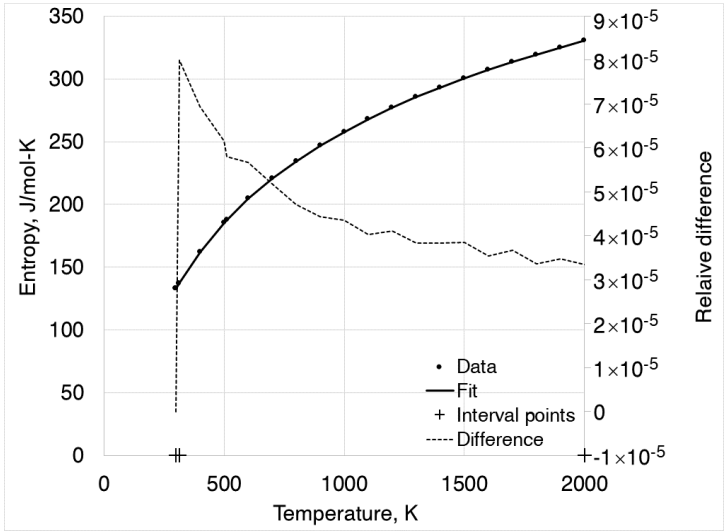


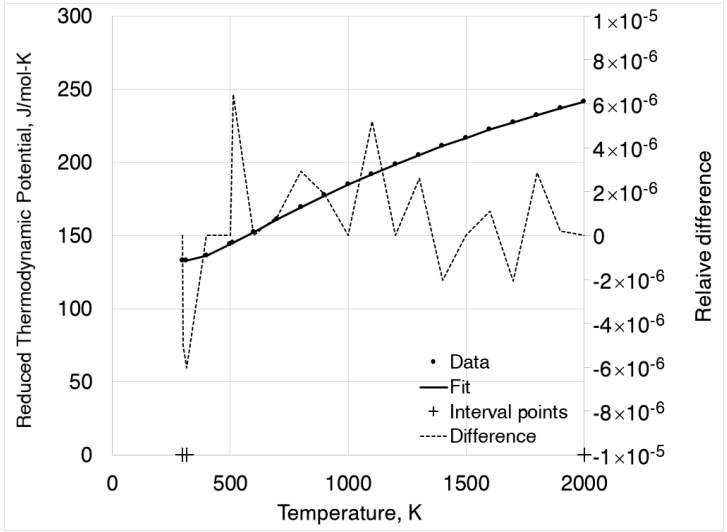
Potassium Fluoride ($K(HF_2)$) – F2H1K1(l)

MW = 78.103 g/mol – $\Delta_f H^o$(298.15 K) = -916.573 kJ/mol – S^o(298.15 K) = 132.764 J/mol-K – p^o = 0.1 MPa



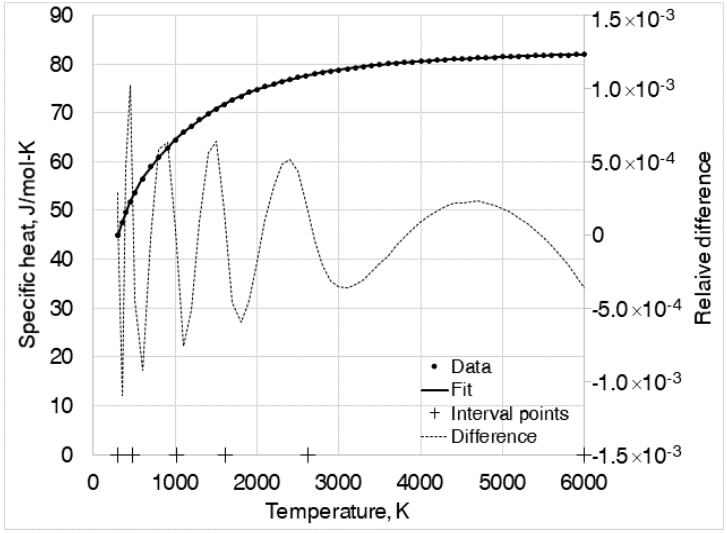


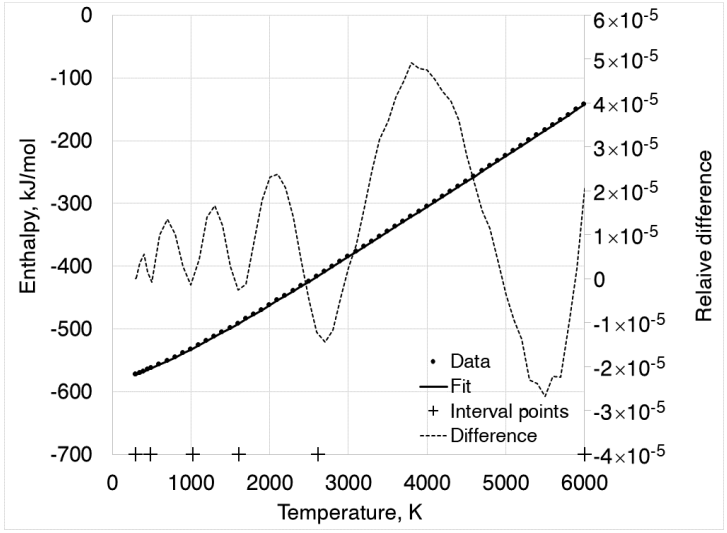


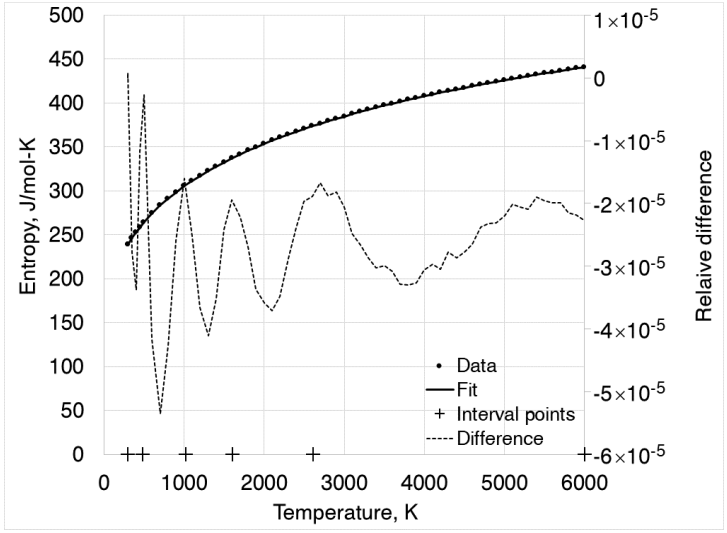


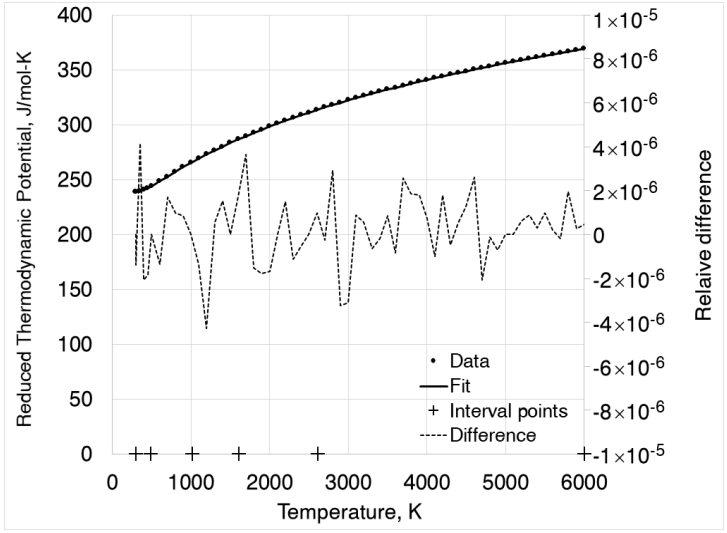
Hydrogen Fluoride ($(HF)_2$) – F2H2(g)

MW = 40.0127 g/mol – $\Delta_f H^o$(298.15 K) = -572.664 kJ/mol – S^o(298.15 K) = 238.849 J/mol-K – p^o = 0.1 MPa



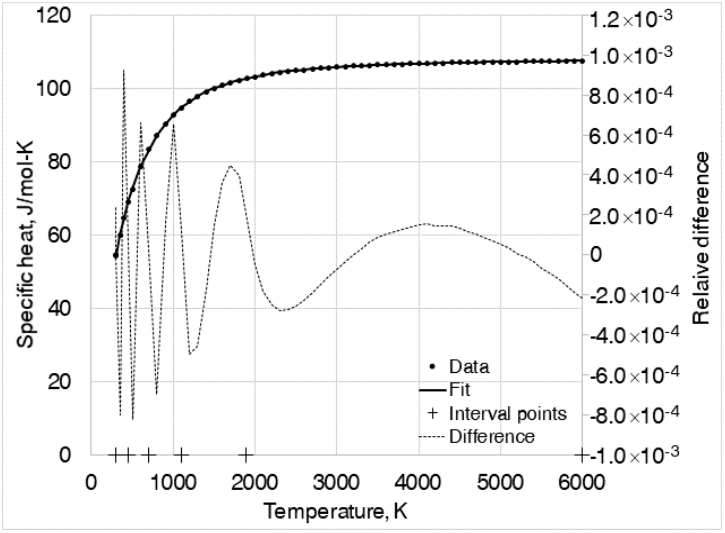


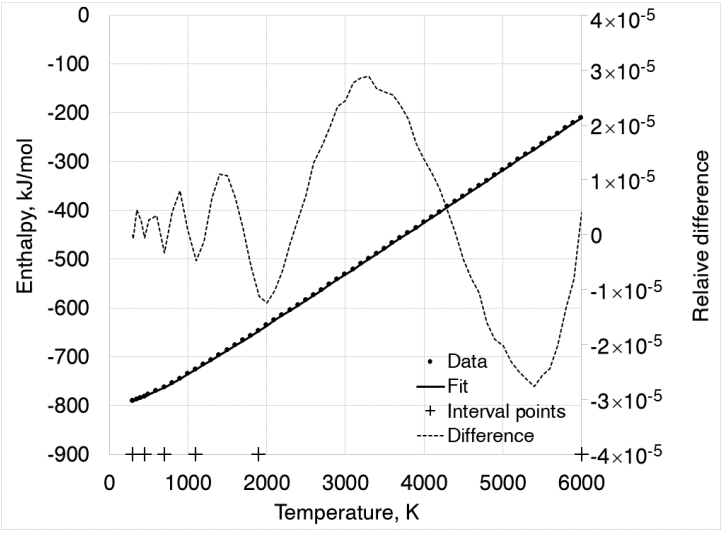


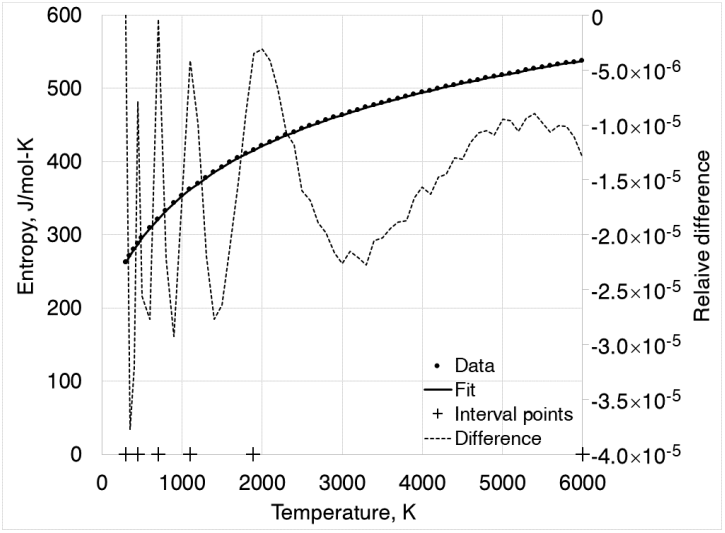


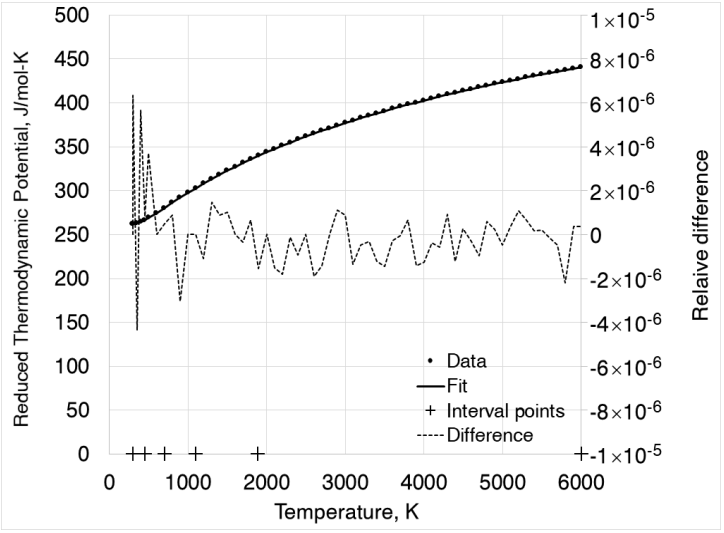
Difluorosilane (SiH_2F_2) – F2H2Si1(g)

MW = 68.0982 g/mol – $\Delta_f H^o$(298.15 K) = -790.776 kJ/mol – S^o(298.15 K) = 262.13 J/mol-K – p^o = 0.1 MPa



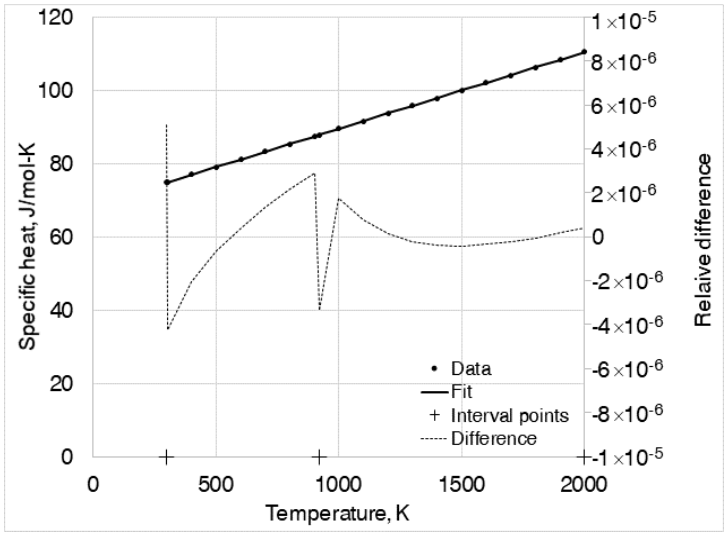


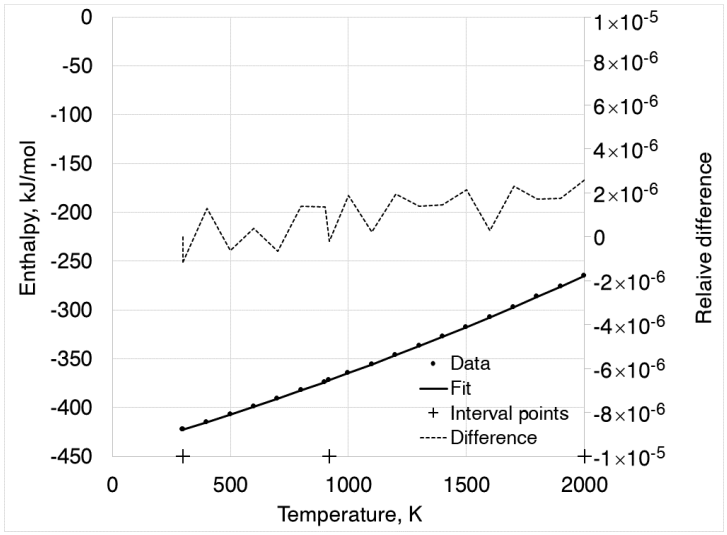


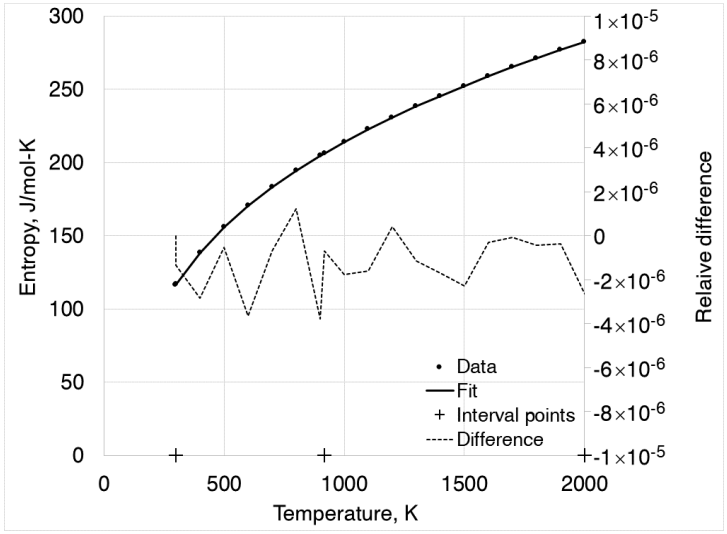


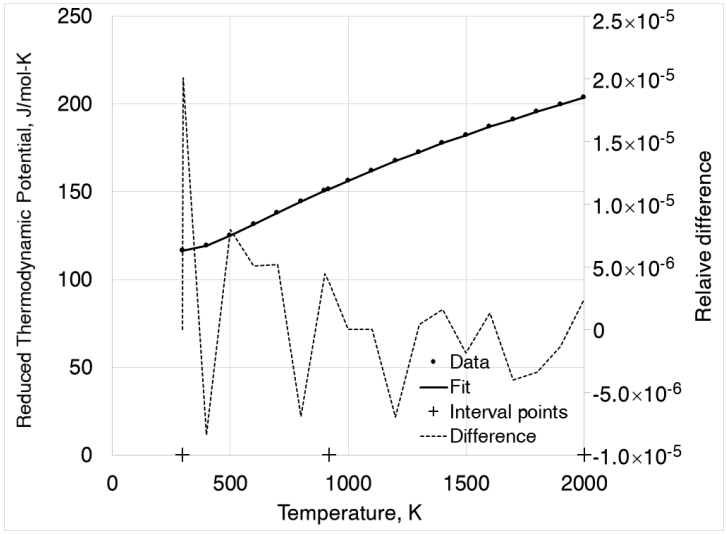
Mercury Fluoride (HgF_2) – F2Hg1(cr)

MW = 238.59 g/mol – $\Delta_f H^o$(298.15 K) = -422.584 kJ/mol – S^o(298.15 K) = 116.315 J/mol-K – p^o = 0.1 MPa



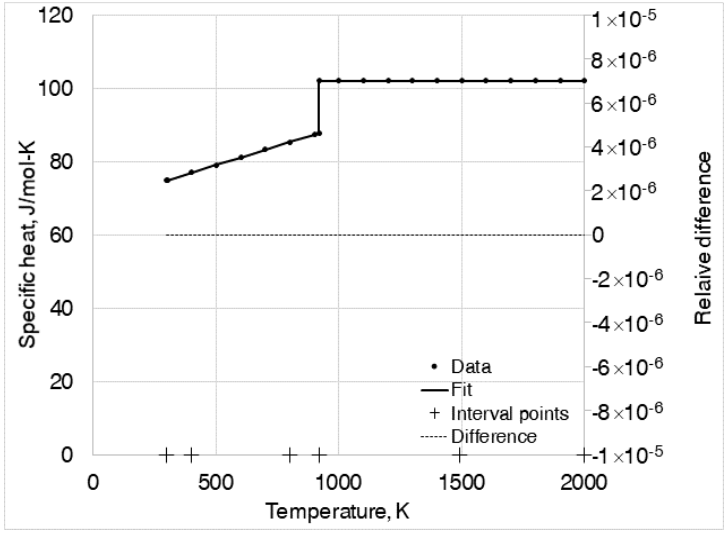


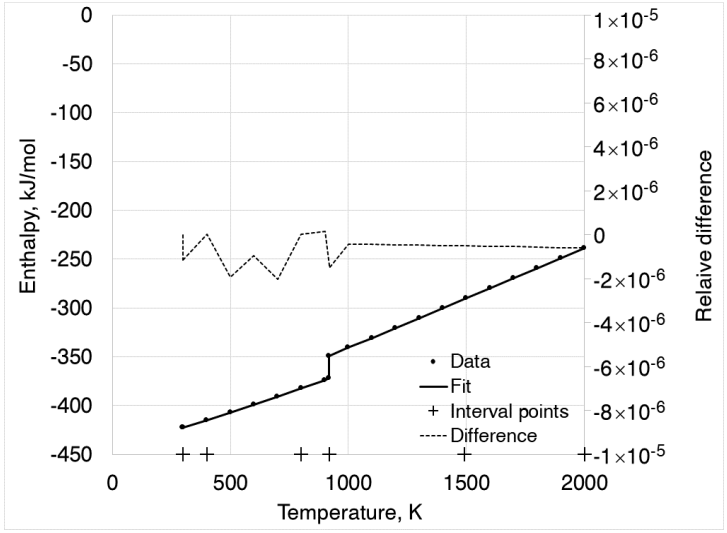


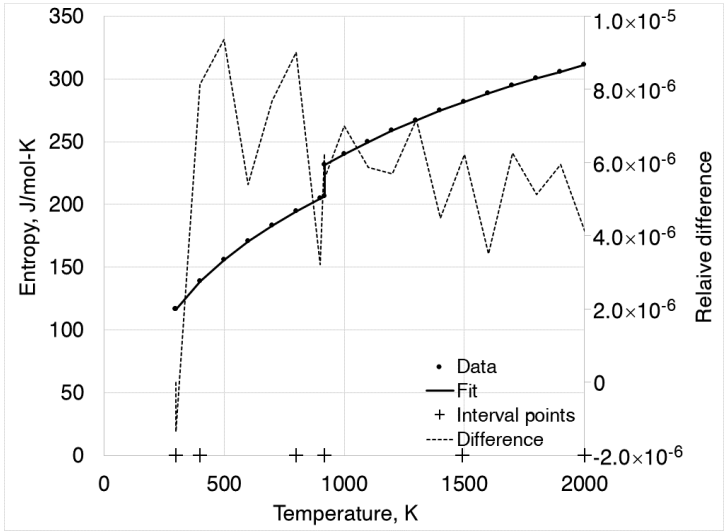


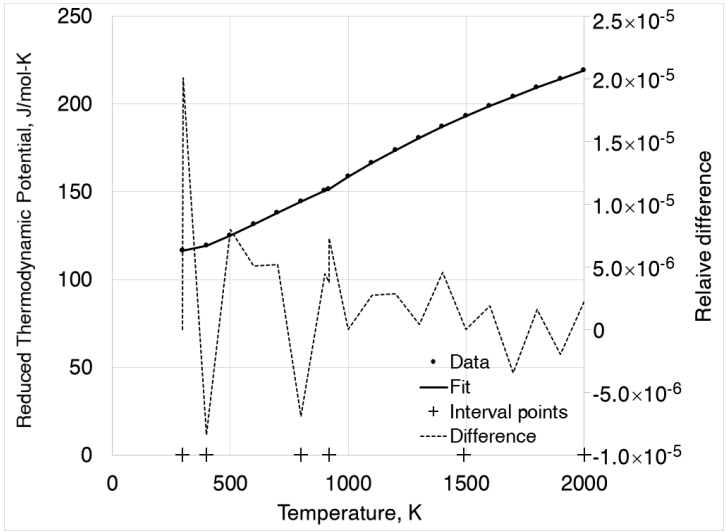
Mercury Fluoride (HgF_2) – F2Hg1(cr,l)

MW = 238.59 g/mol – $\Delta_f H^o$(298.15 K) = -422.584 kJ/mol – S^o(298.15 K) = 116.315 J/mol-K – p^o = 0.1 MPa



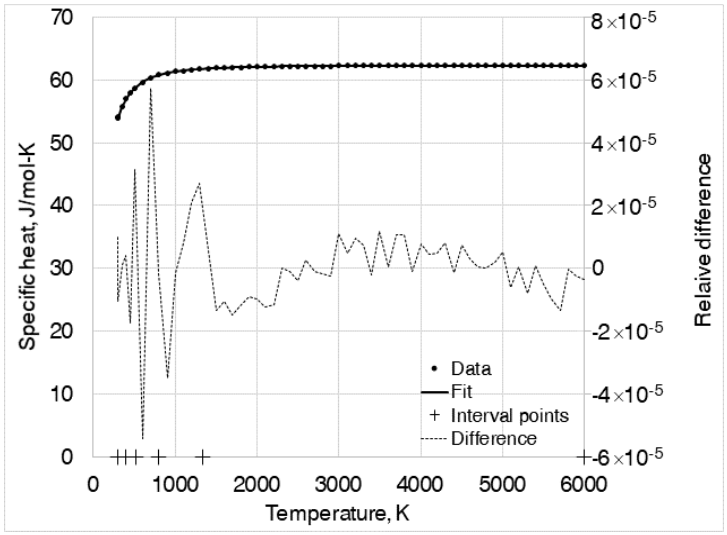


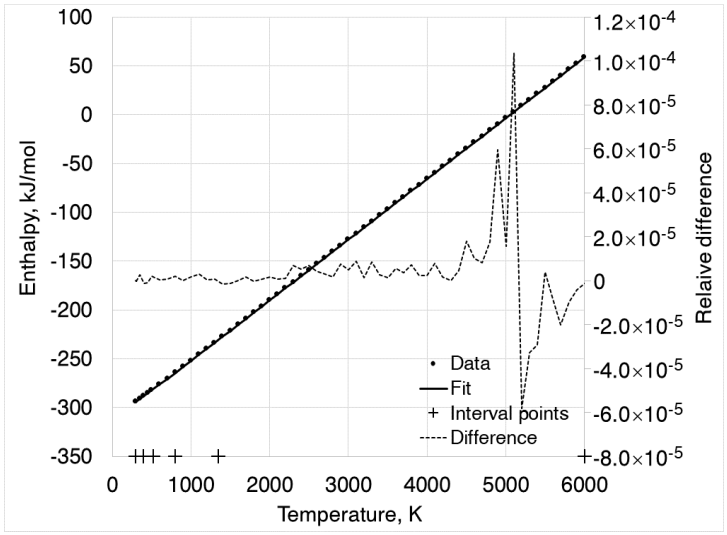


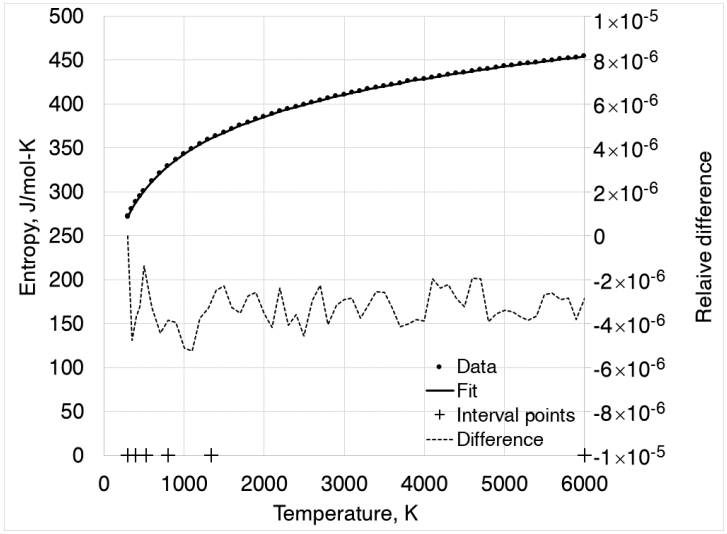


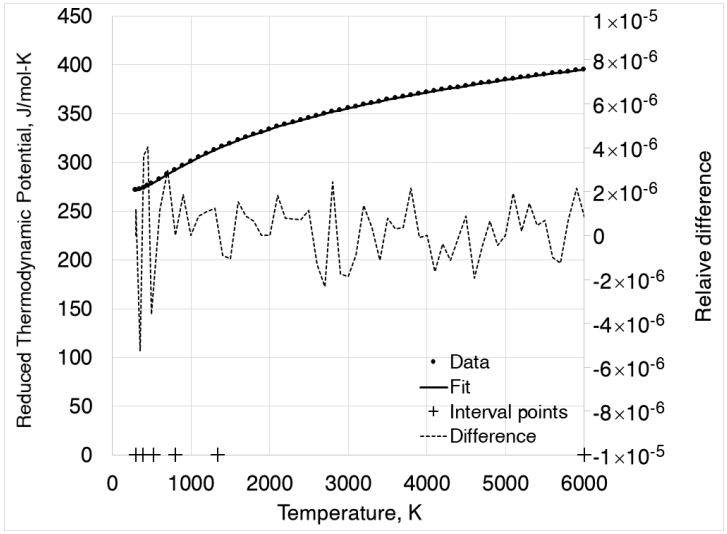
Mercury Fluoride (HgF_2) – F2Hg1(g)

MW = 238.59 g/mol – $\Delta_f H^o$(298.15 K) = -293.654 kJ/mol – S^o(298.15 K) = 271.737 J/mol-K – p^o = 0.1 MPa



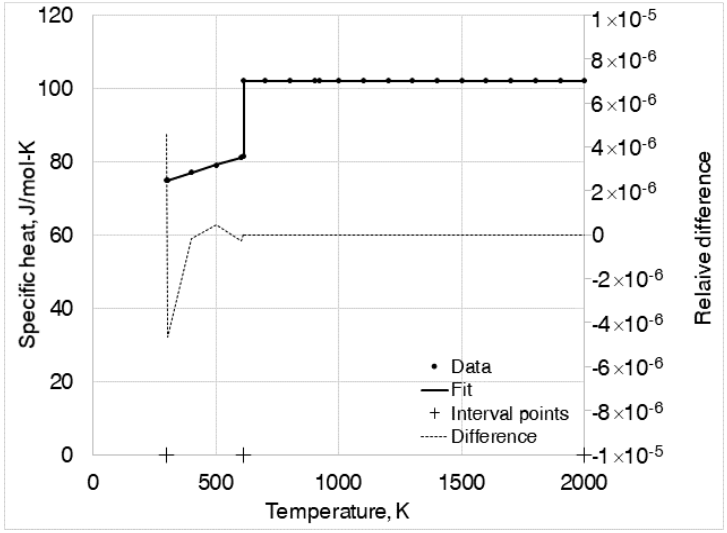


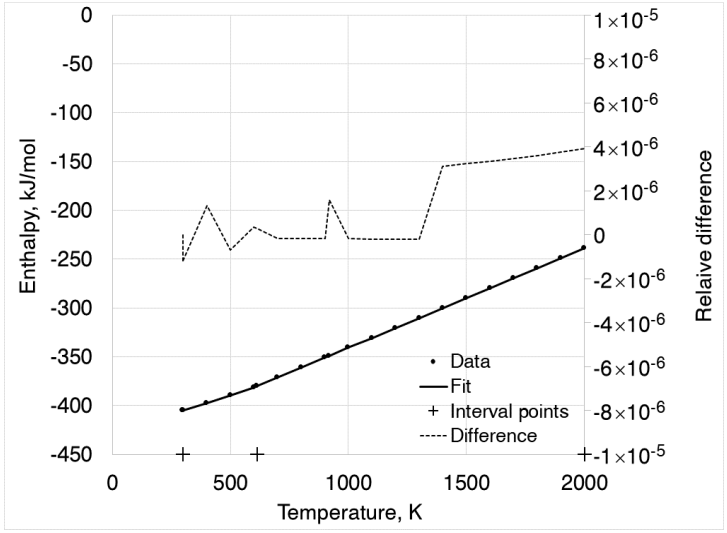


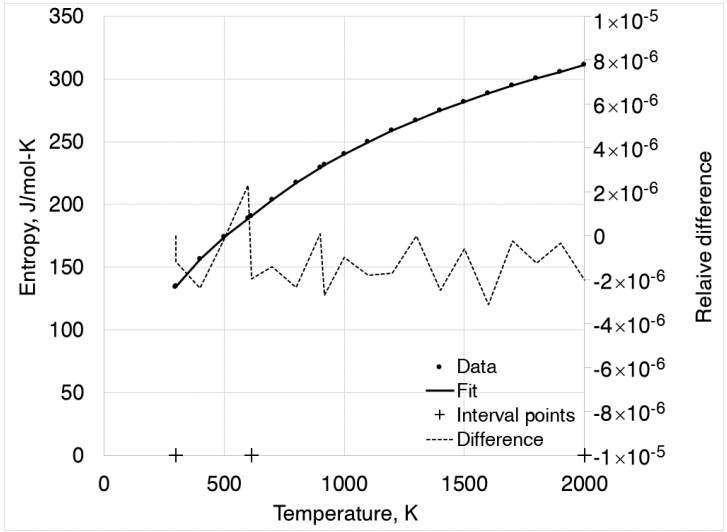


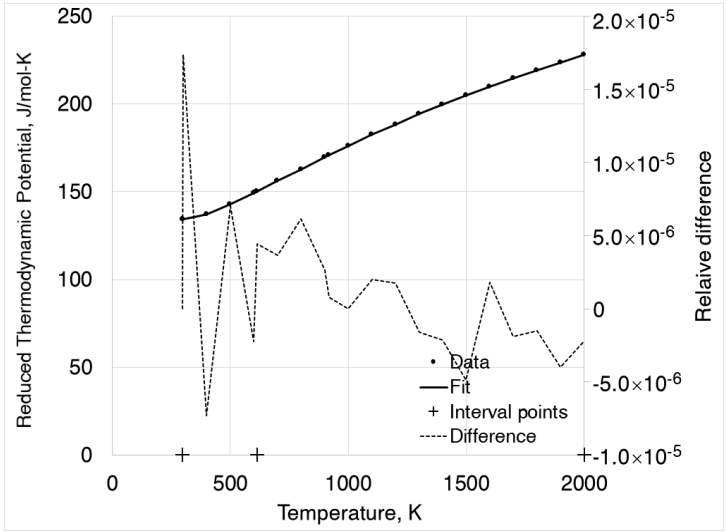
Mercury Fluoride (HgF_2) – F2Hg1(l)

MW = 238.59 g/mol – $\Delta_f H^o$(298.15 K) = -404.897 kJ/mol – S^o(298.15 K) = 134.246 J/mol-K – p^o = 0.1 MPa



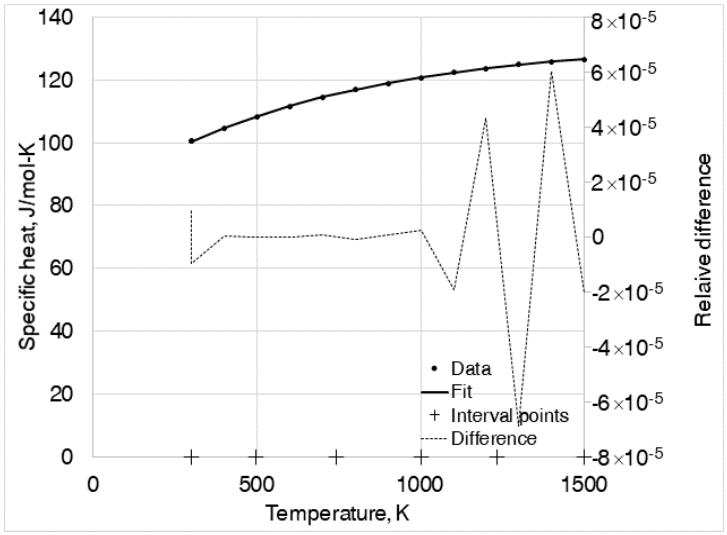


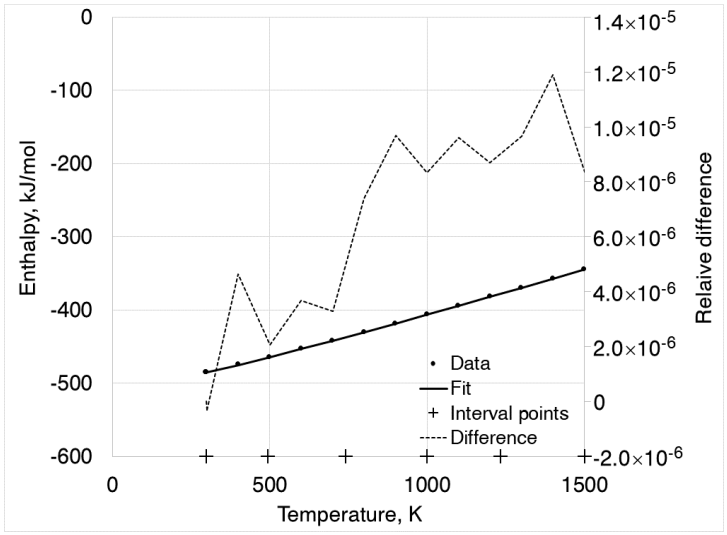


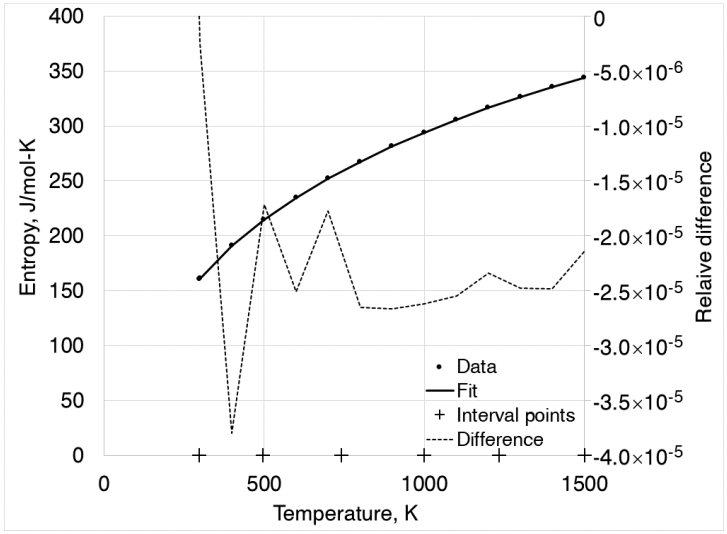


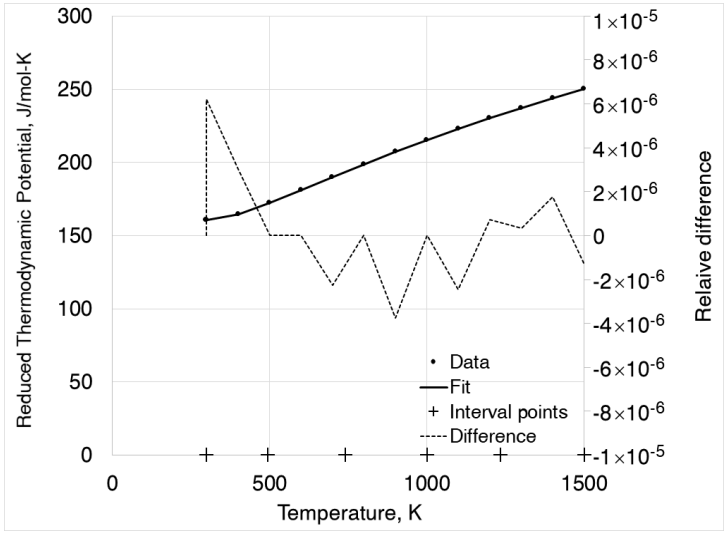
Mercury Fluoride (Hg_2F_2) – F2Hg2(cr)

MW = 439.18 g/mol – $\Delta_f H^o$(298.15 K) = -485.344 kJ/mol – S^o(298.15 K) = 160.666 J/mol-K – p^o = 0.1 MPa



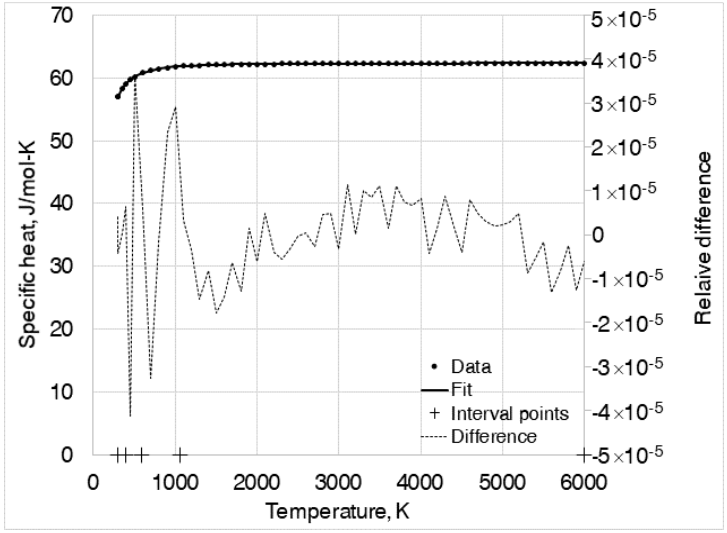


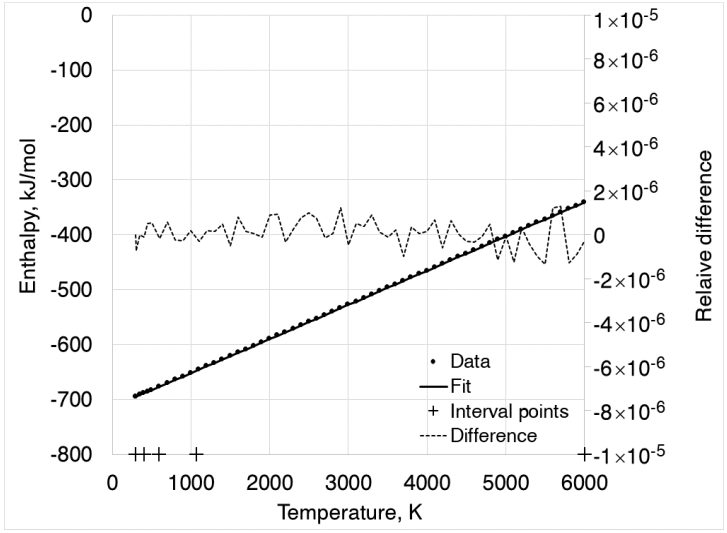


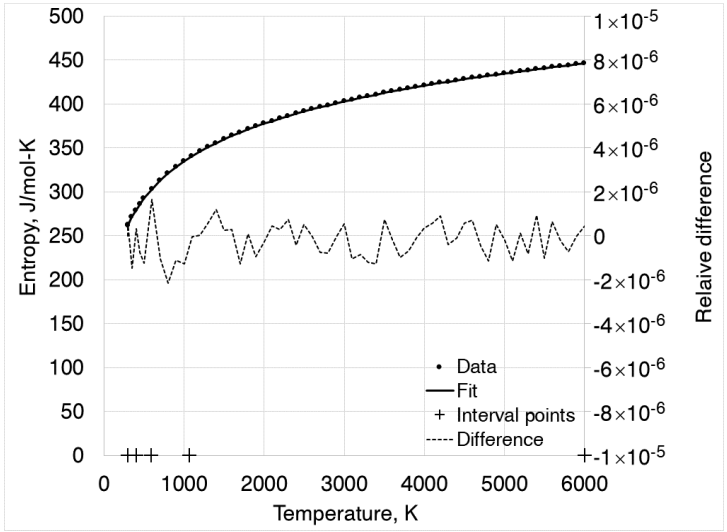


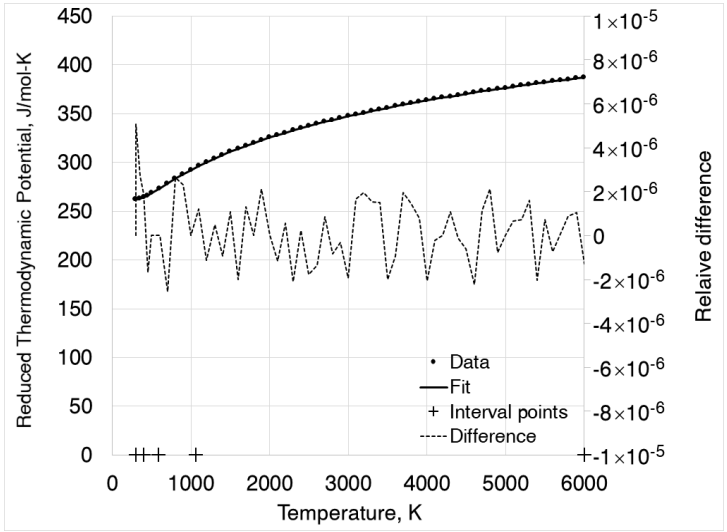
Potassium Fluoride, Ion (KF_2^-) – F2K1-(g)

MW = 77.0957 g/mol – $\Delta_f H^o$(298.15 K) = -694.544 kJ/mol – S^o(298.15 K) = 261.868 J/mol-K – p^o = 0.1 MPa



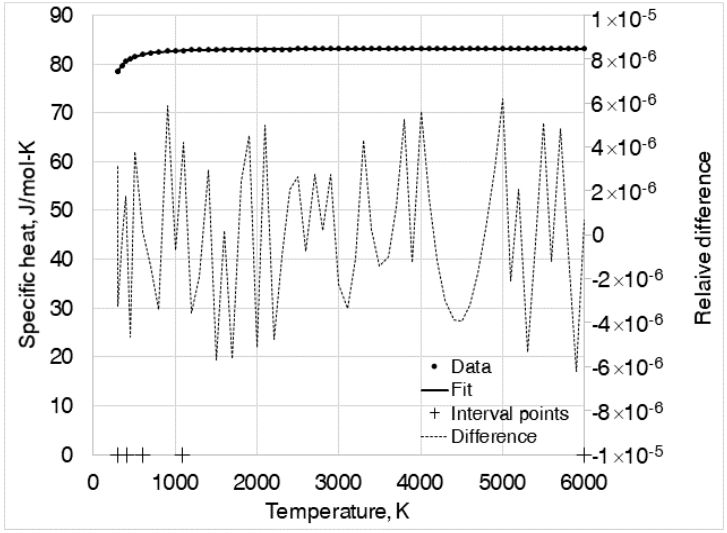


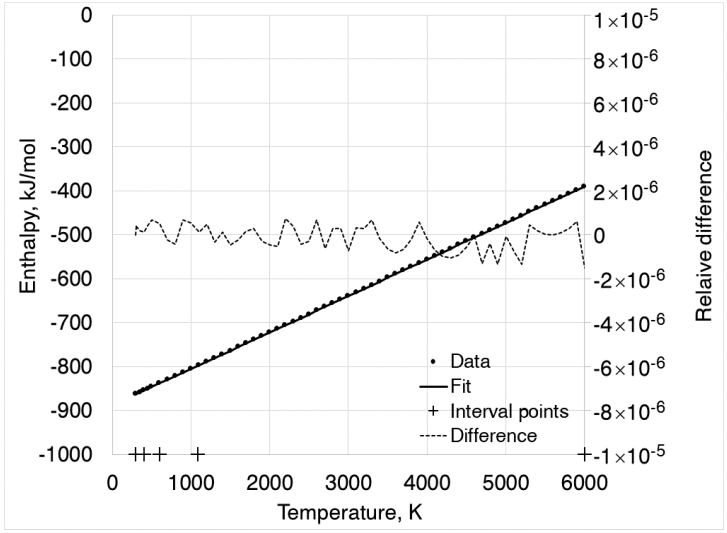


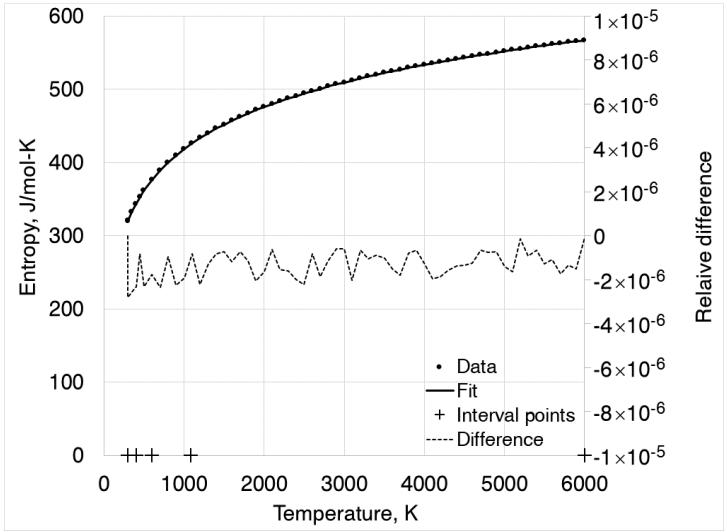


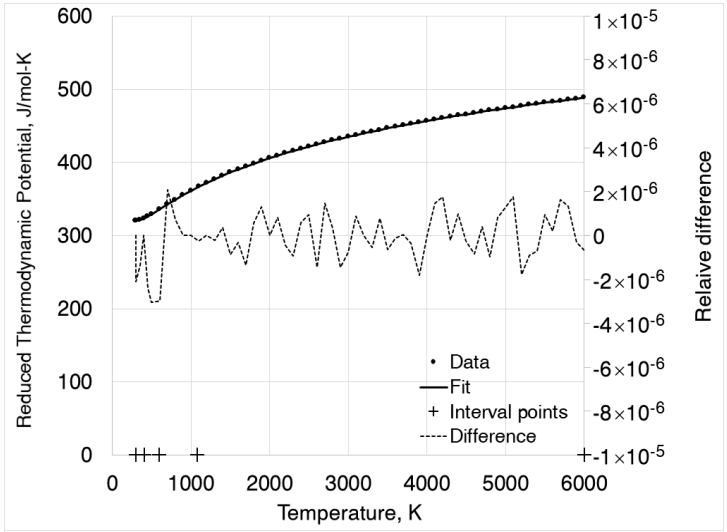
Potassium Fluoride ($(KF)_2$) – F2K2(g)

MW = 116.1934 g/mol – $\Delta_f H^o$(298.15 K) = -862.741 kJ/mol – S^o(298.15 K) = 319.955 J/mol-K – p^o = 0.1 MPa



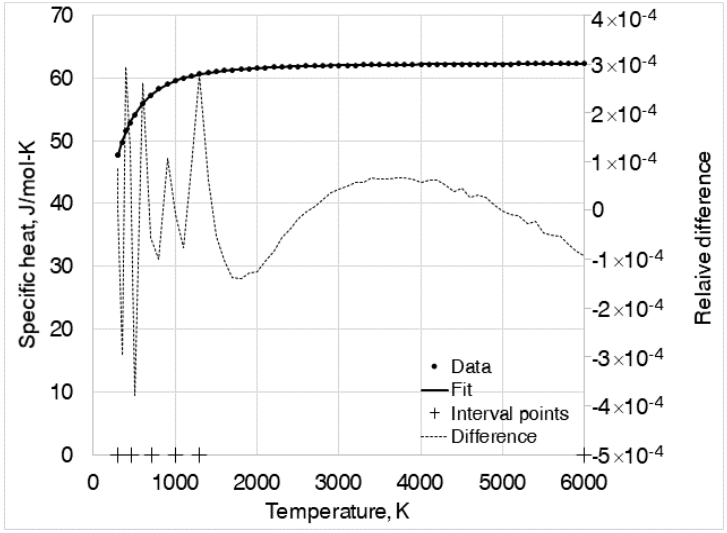


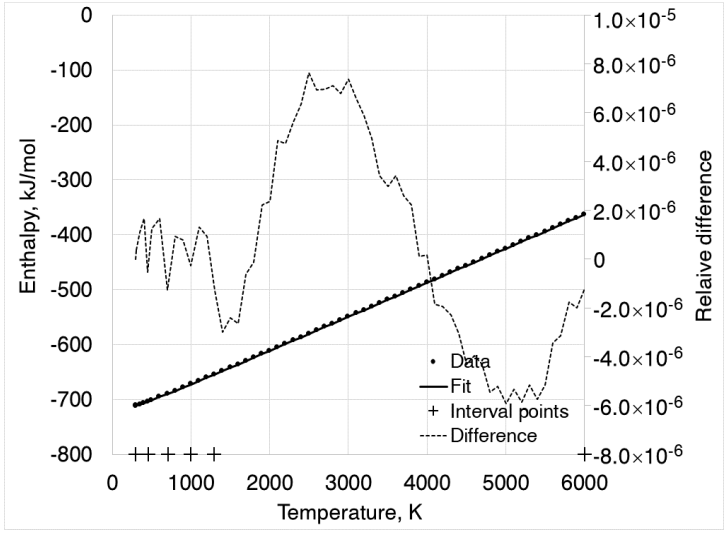


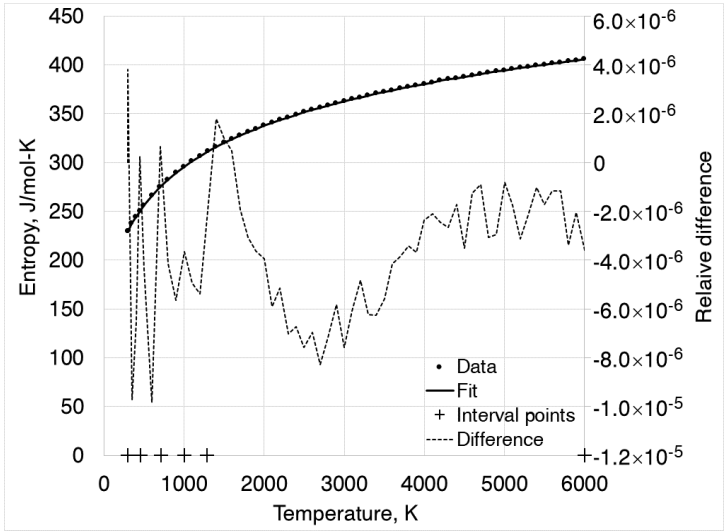


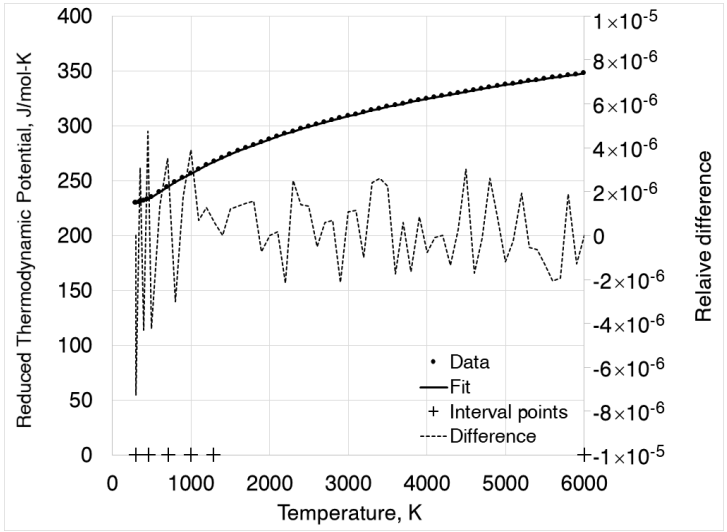
Lithium Fluoride, Ion (LiF_2^-) – F2Li1-(g)

MW = 44.938 g/mol – $\Delta_f H^o$(298.15 K) = -711.28 kJ/mol – S^o(298.15 K) = 229.618 J/mol-K – p^o = 0.1 MPa



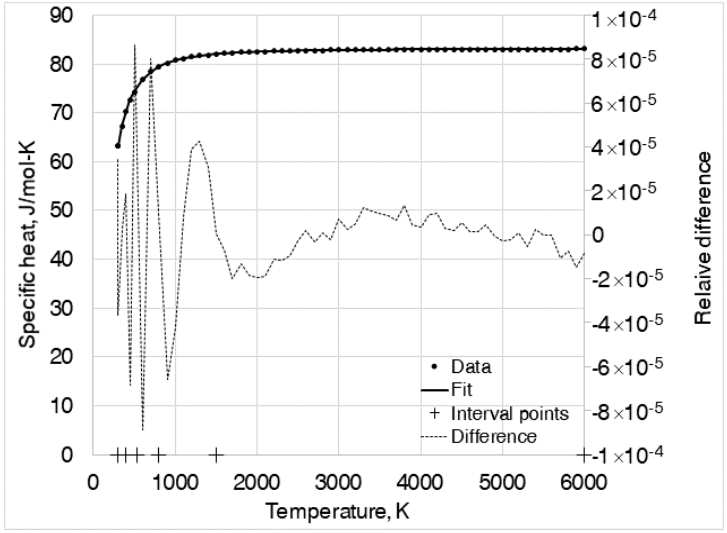


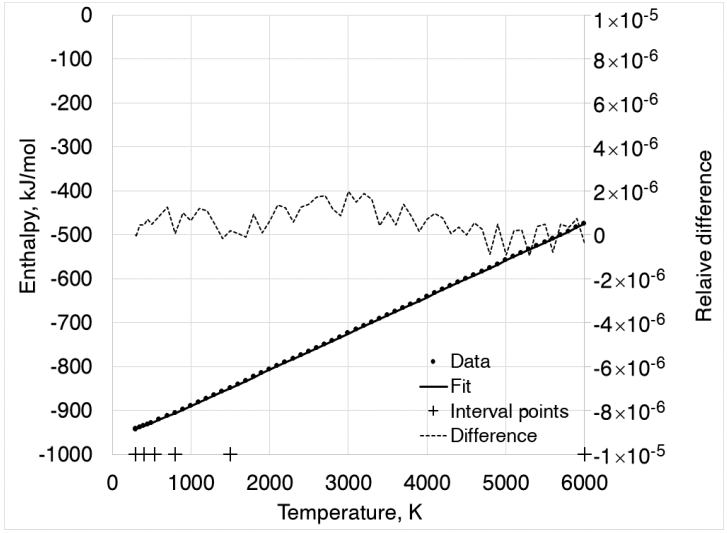


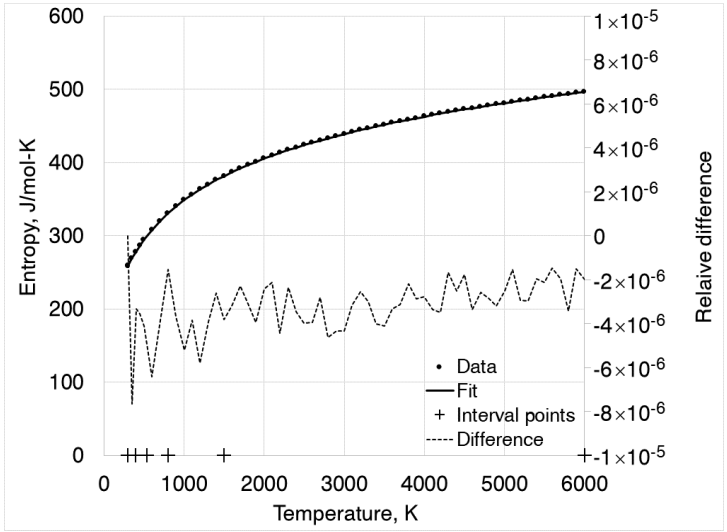


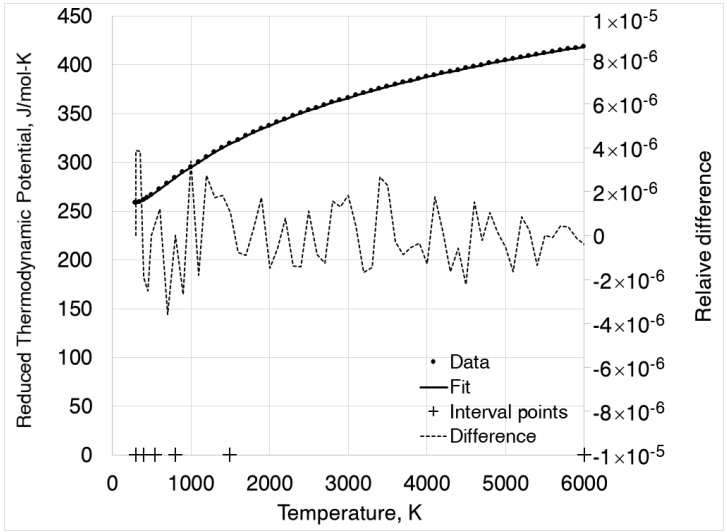
Lithium Fluoride ($(LiF)_2$) – F2Li2(g)

MW = 51.879 g/mol – $\Delta_f H^o$(298.15 K) = -942.781 kJ/mol – S^o(298.15 K) = 258.627 J/mol-K – p^o = 0.1 MPa



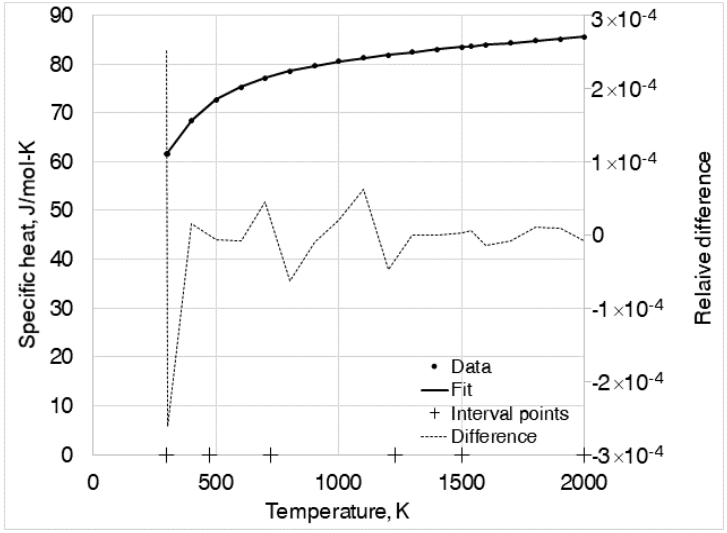


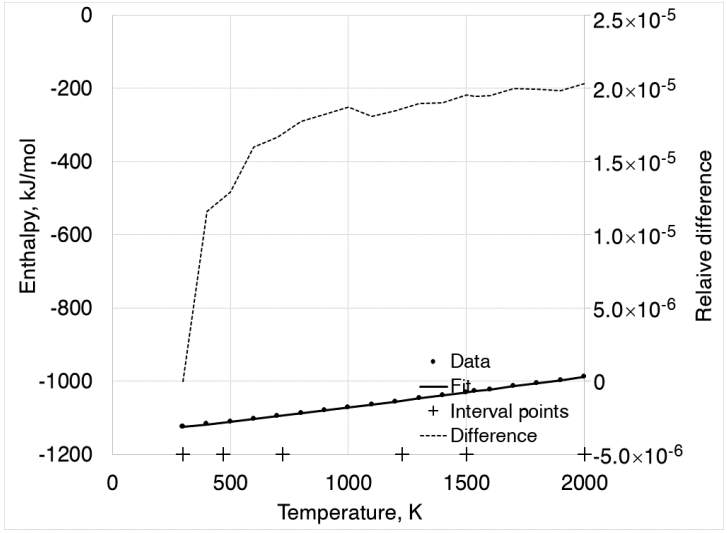


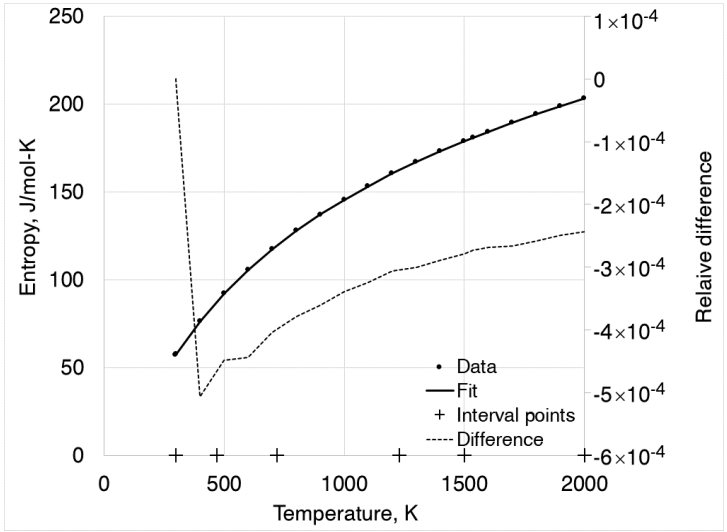


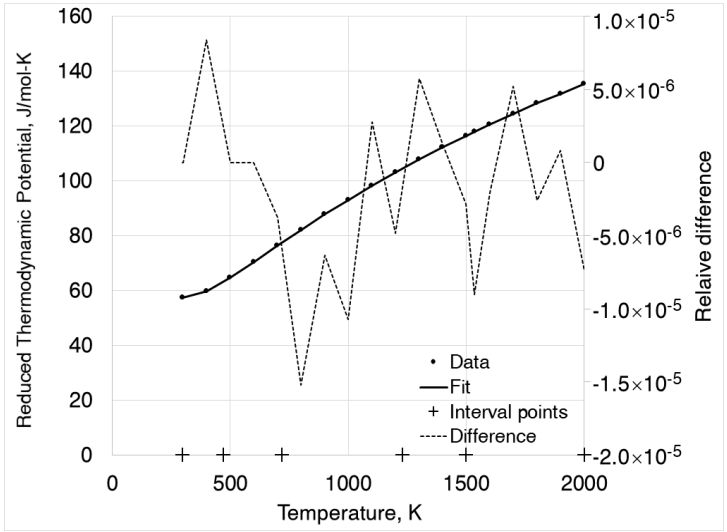
Magnesium Fluoride (MgF_2) – F2Mg1(cr)

MW = 62.302 g/mol – $\Delta_f H^o$(298.15 K) = -1124.241 kJ/mol – S^o(298.15 K) = 57.255 J/mol-K – p^o = 0.1 MPa



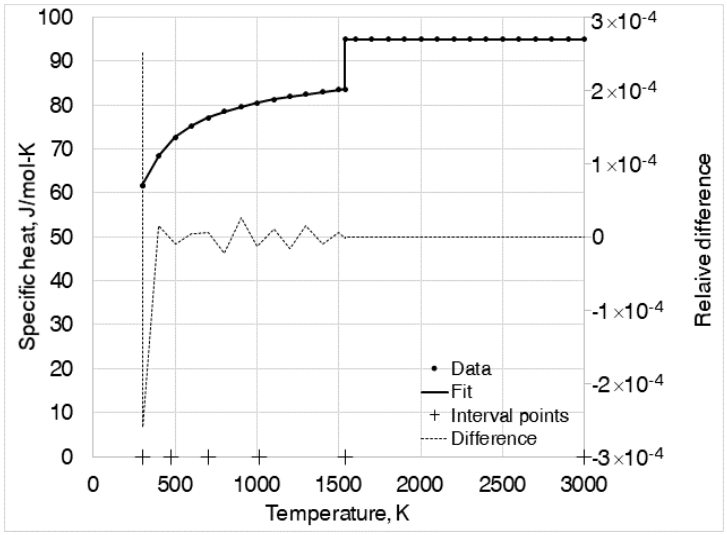


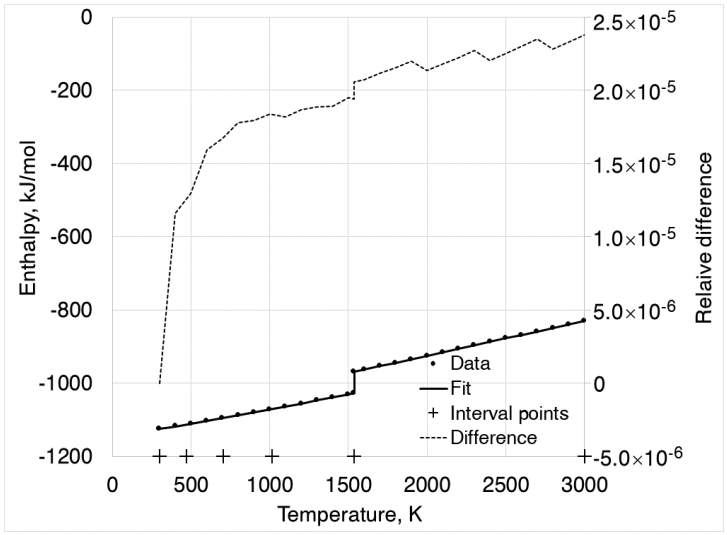


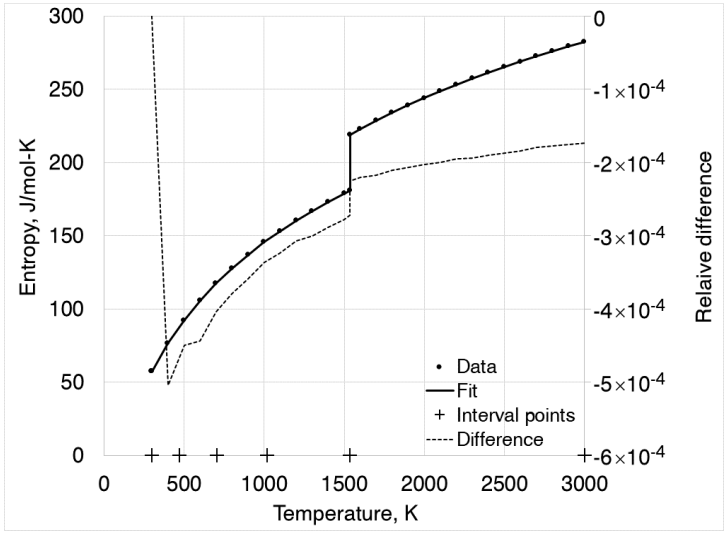


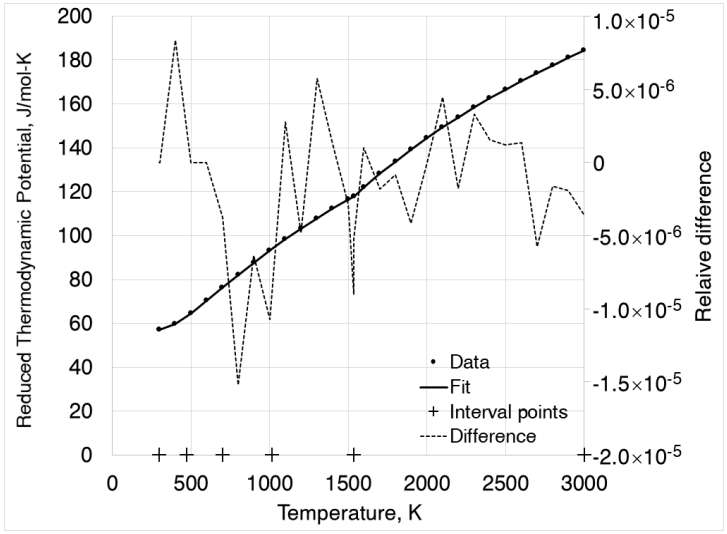
Magnesium Fluoride (MgF_2) – F2Mg1(cr,l)

MW = 62.302 g/mol – $\Delta_f H^o$(298.15 K) = -1124.241 kJ/mol – S^o(298.15 K) = 57.255 J/mol-K – p^o = 0.1 MPa



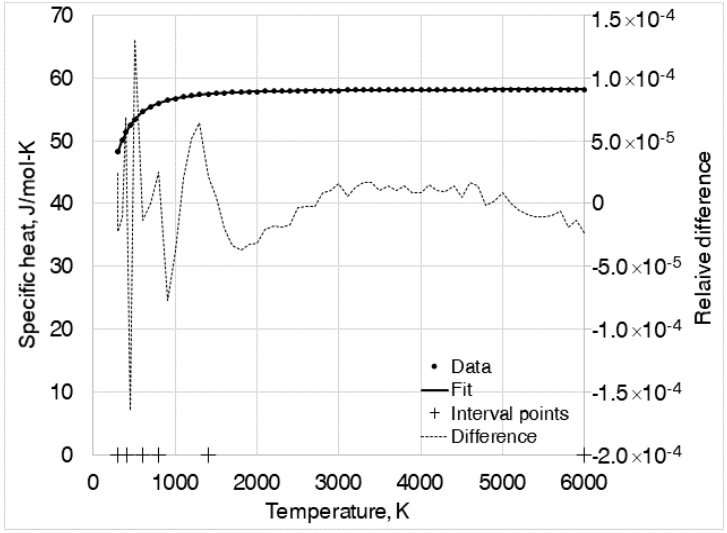


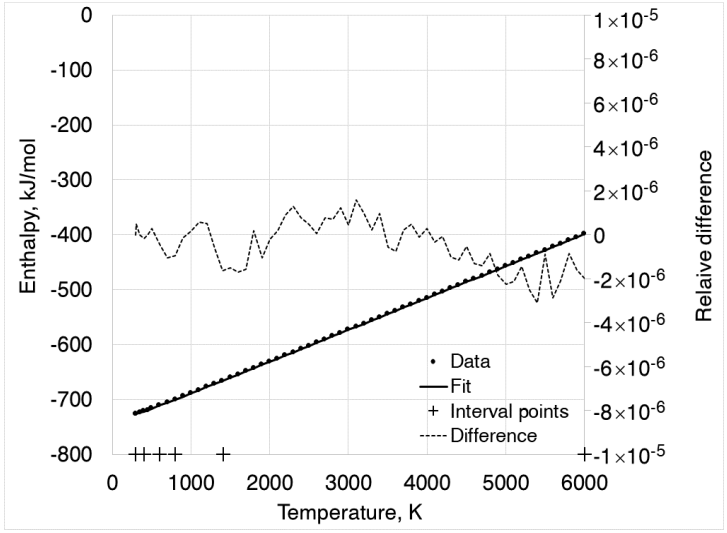


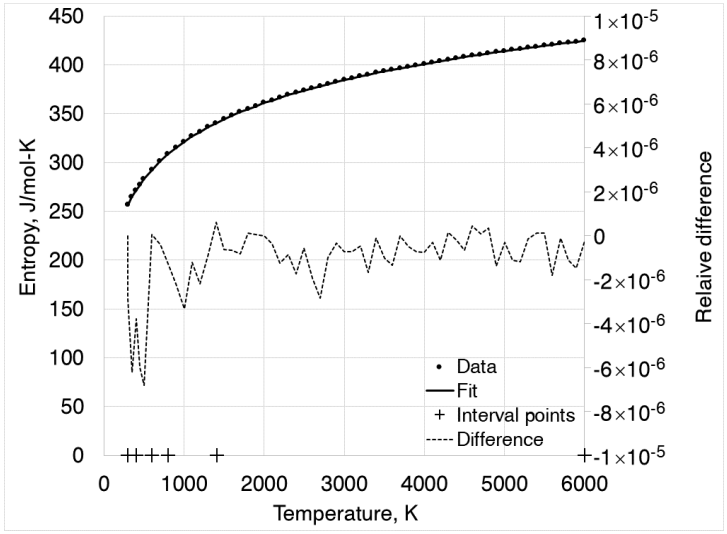


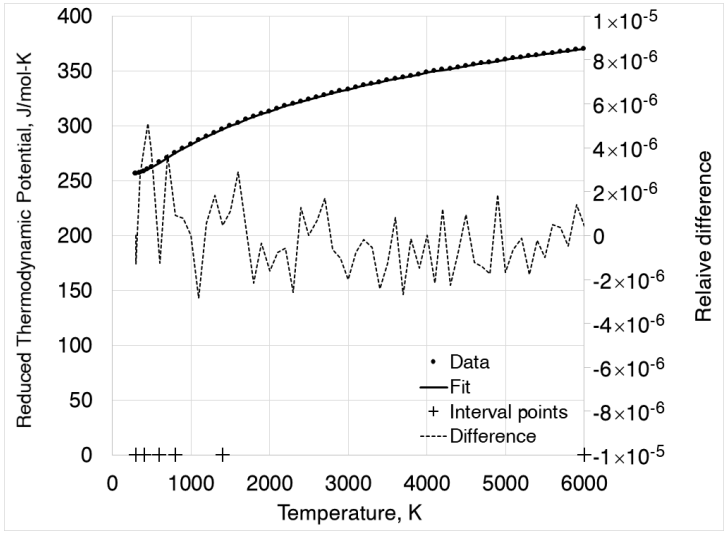
Magnesium Fluoride (MgF_2) – F2Mg1(g)

MW = 62.302 g/mol – $\Delta_f H^o$(298.15 K) = -726.761 kJ/mol – S^o(298.15 K) = 256.508 J/mol-K – p^o = 0.1 MPa



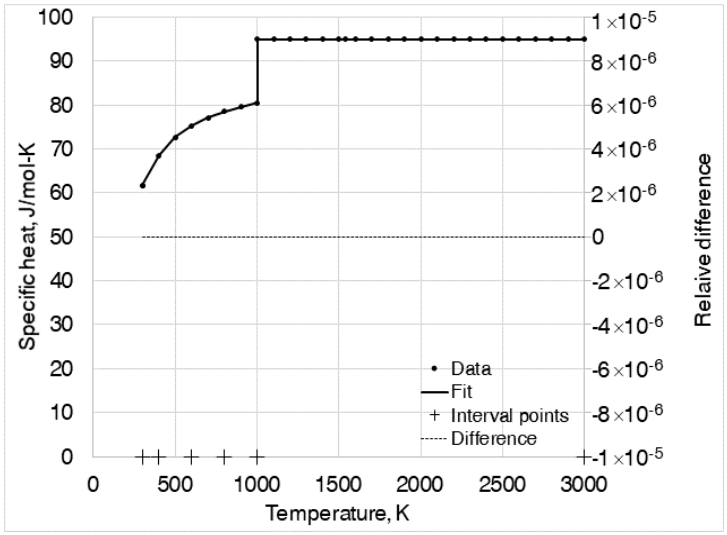


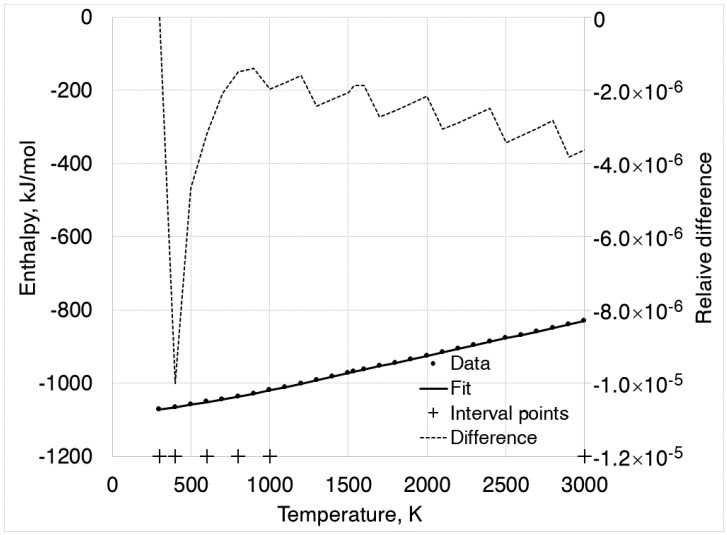


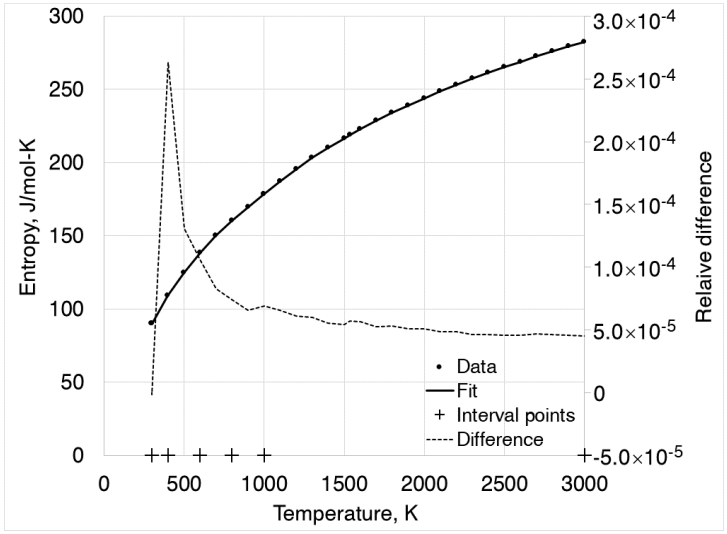


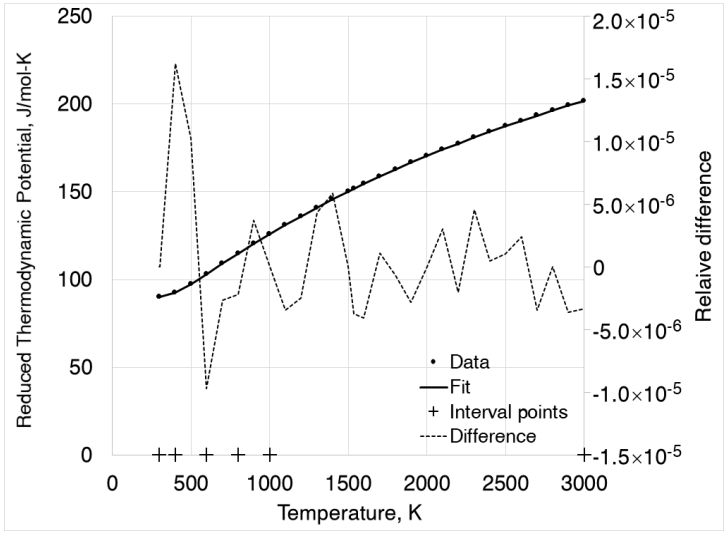
Magnesium Fluoride (MgF_2) – F2Mg1(l)

MW = 62.302 g/mol – $\Delta_f H^o$(298.15 K) = -1072.356 kJ/mol – S^o(298.15 K) = 89.967 J/mol-K – p^o = 0.1 MPa



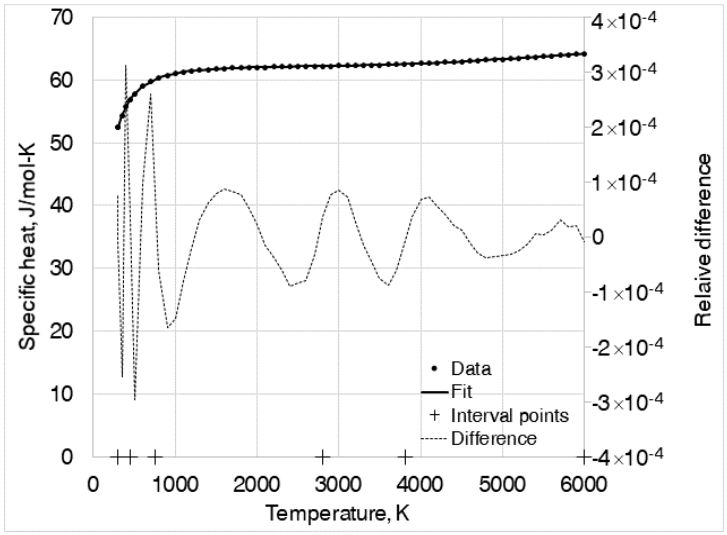


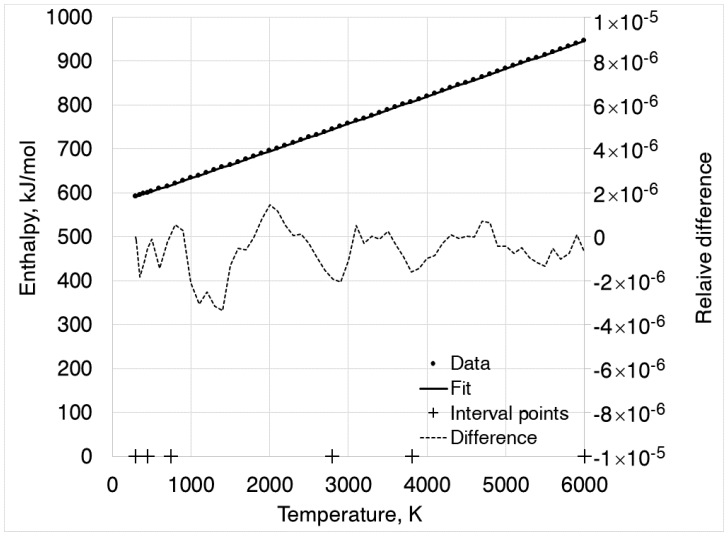


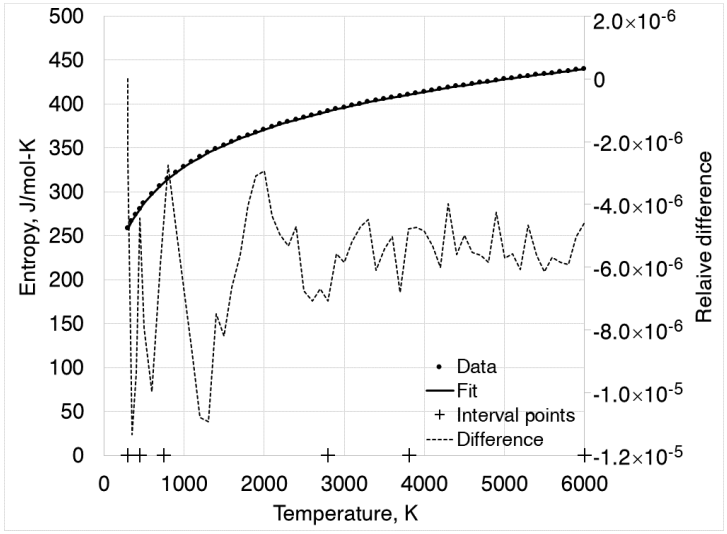


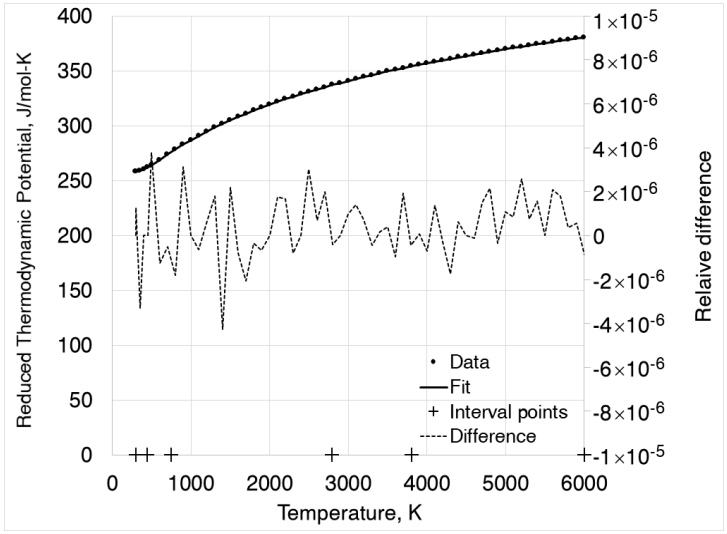
Magnesium Fluoride, Ion (MgF_2^+) – F2Mg1+(g)

MW = 62.301 g/mol – $\Delta_f H^o$(298.15 K) = 591.982 kJ/mol – S^o(298.15 K) = 258.145 J/mol-K – p^o = 0.1 MPa



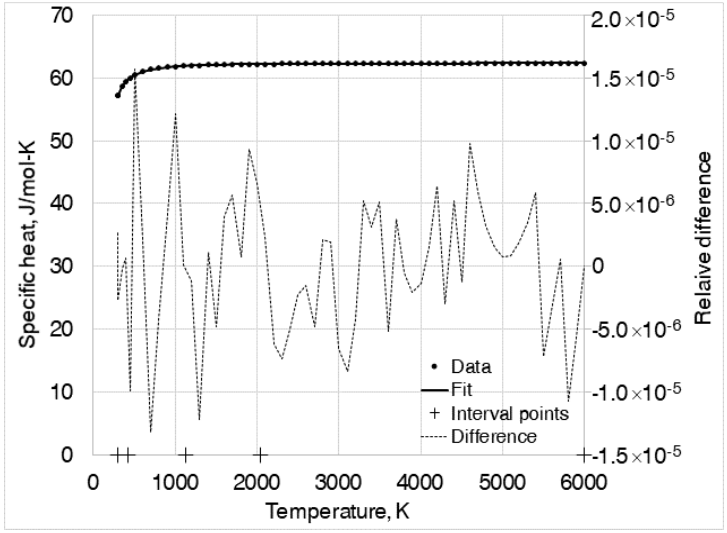


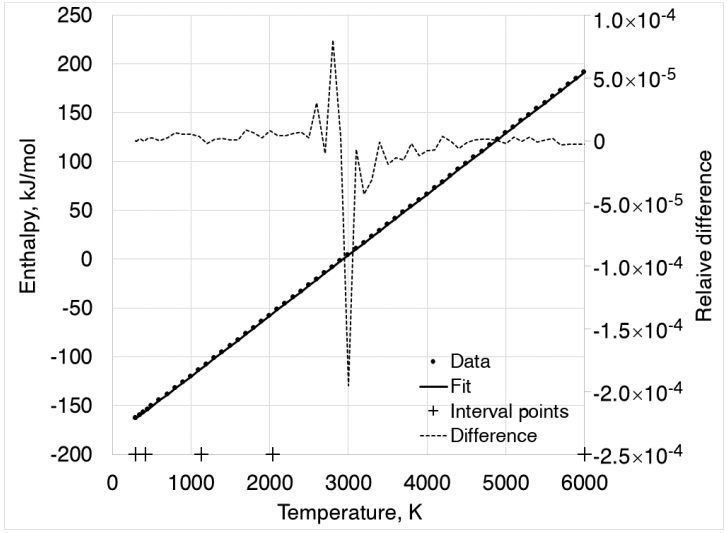


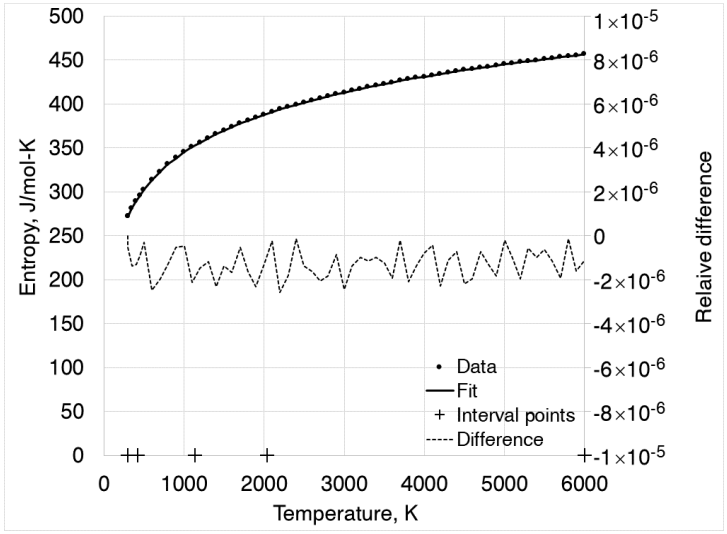


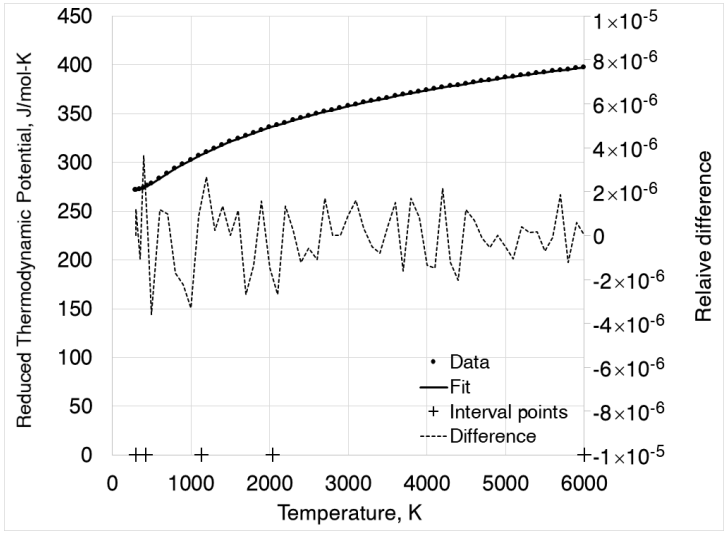
Molybdenum Fluoride (MoF_2) – F2Mo1(g)

MW = 133.94 g/mol – $\Delta_f H^o$(298.15 K) = -162.758 kJ/mol – S^o(298.15 K) = 271.932 J/mol-K – p^o = 0.1 MPa



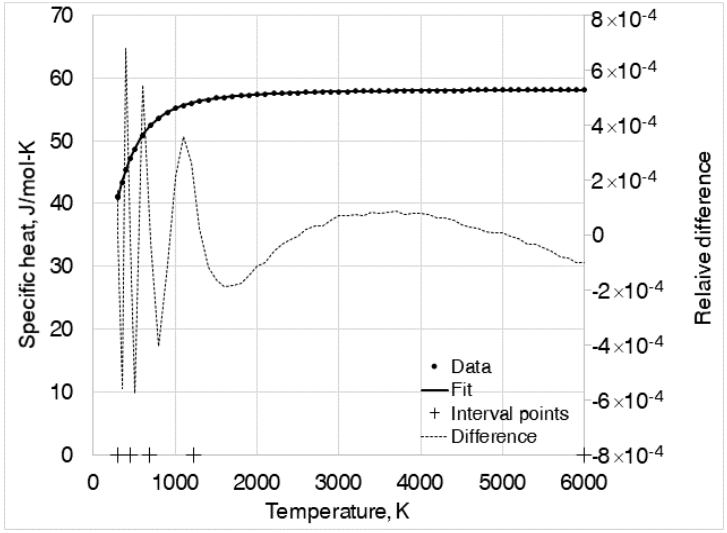


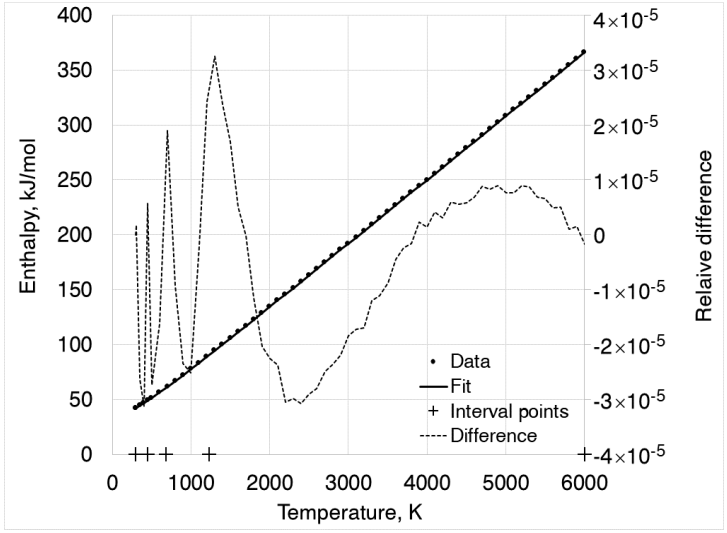


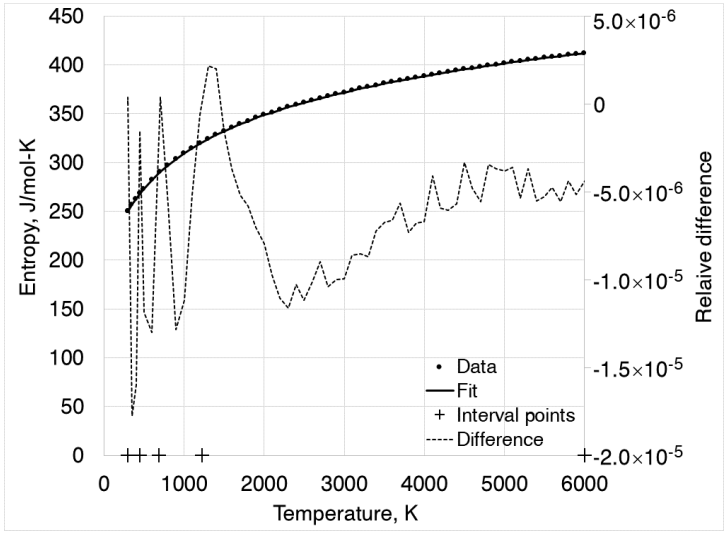


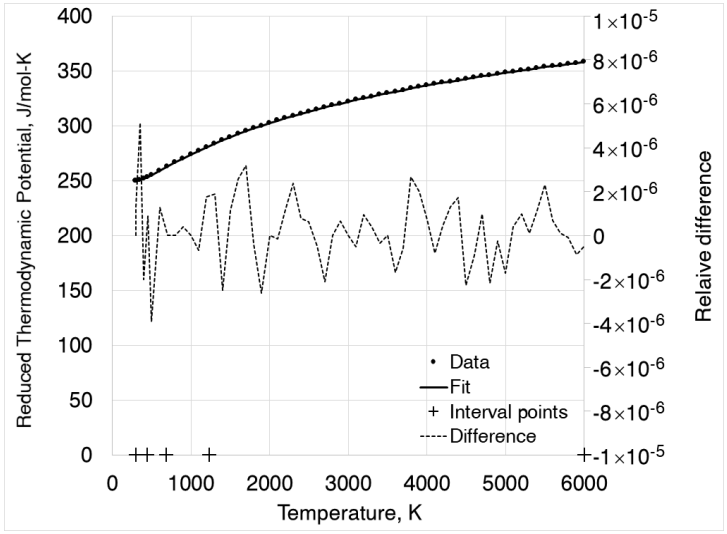
Difluoroamidogen (NF_2) – F2N1(g)

MW = 52.0035 g/mol – $\Delta_f H^o$(298.15 K) = 42.258 kJ/mol – S^o(298.15 K) = 249.95 J/mol-K – p^o = 0.1 MPa



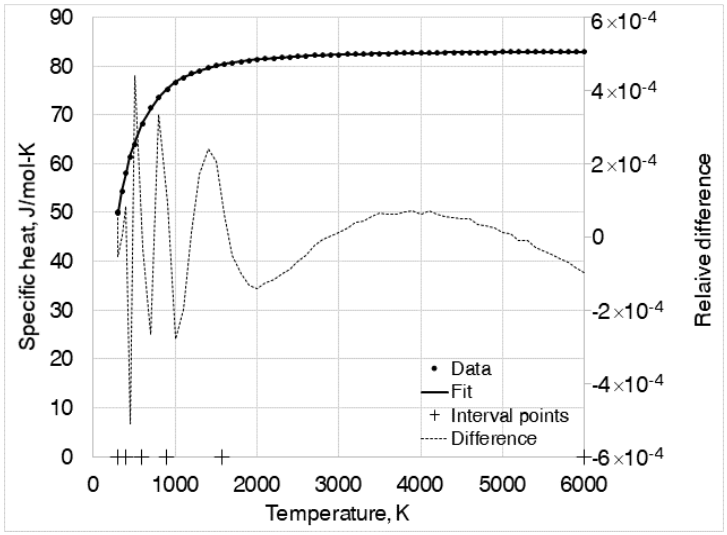


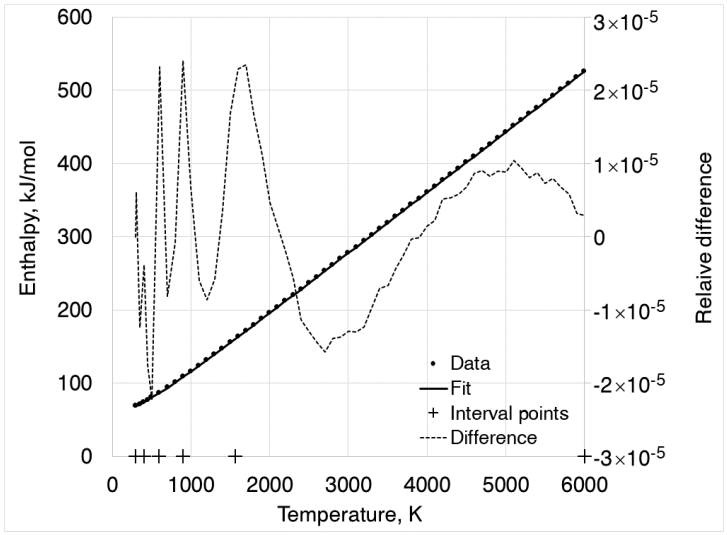


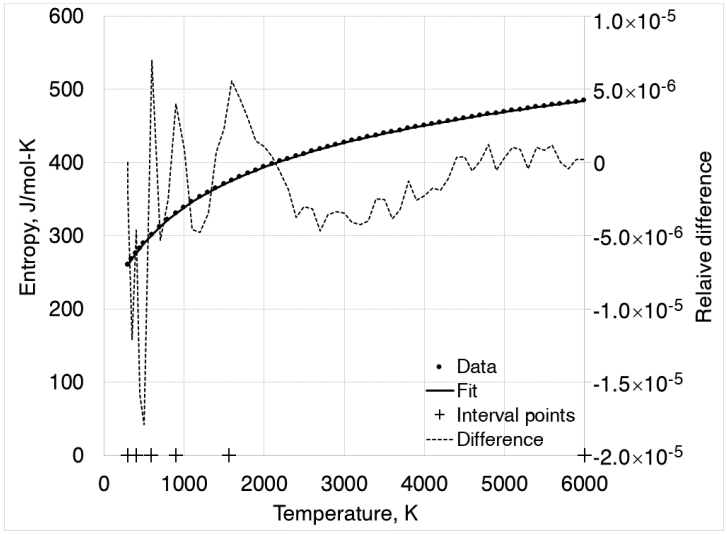


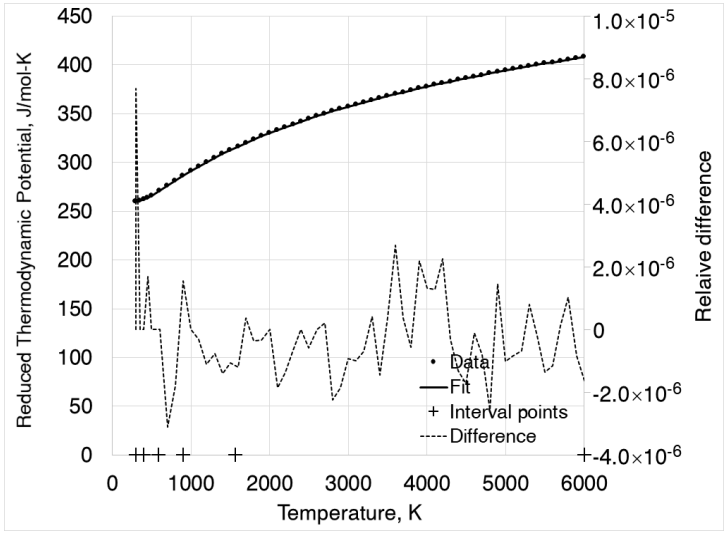
Nitrogen Fluoride, Cis (FNNF) – F2N2(g)_01

MW = 66.0102 g/mol – $\Delta_f H^o$(298.15 K) = 68.618 kJ/mol – S^o(298.15 K) = 259.81 J/mol-K – p^o = 0.1 MPa



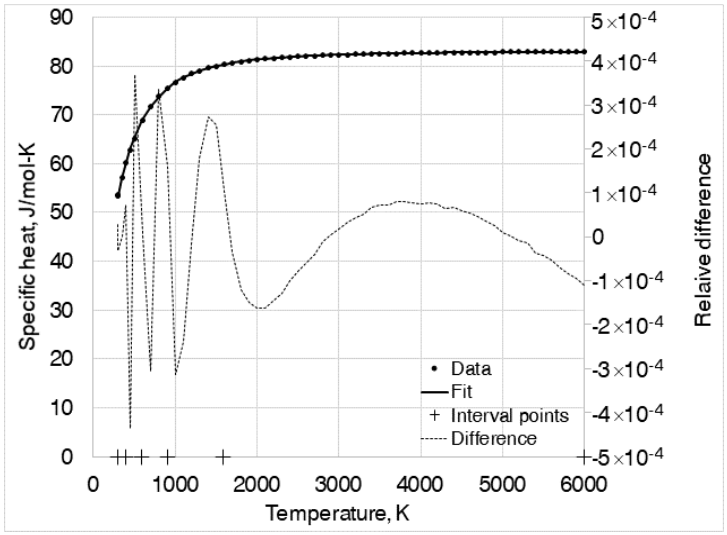


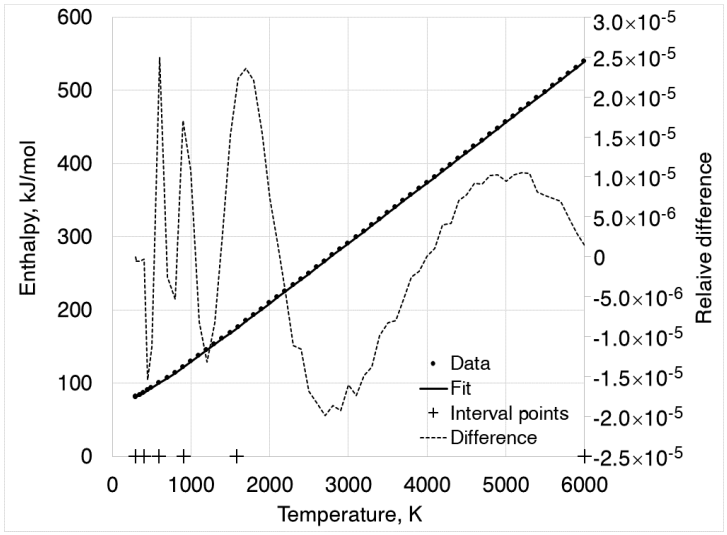


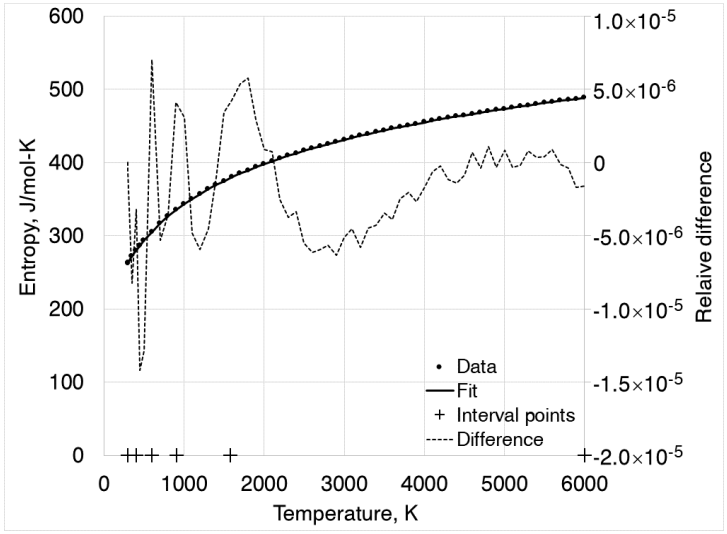


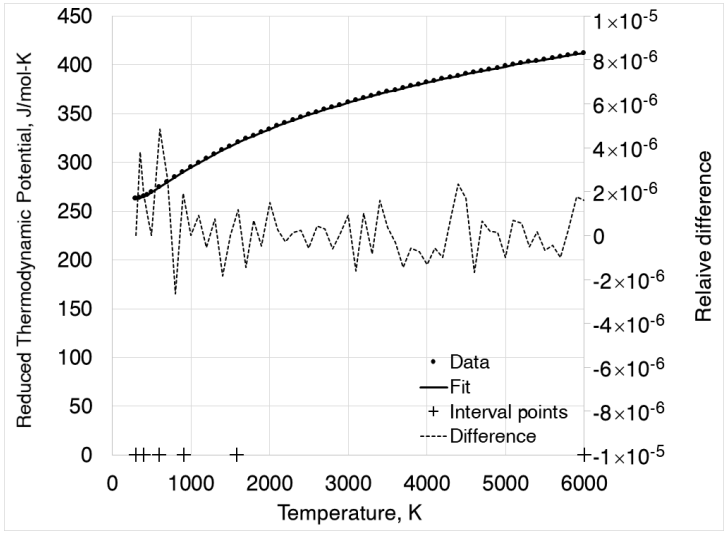
Nitrogen Fluoride, Trans (FNNF) – F2N2(g)_02

MW = 66.0102 g/mol – $\Delta_f H^o$(298.15 K) = 81.17 kJ/mol – S^o(298.15 K) = 262.66 J/mol-K – p^o = 0.1 MPa



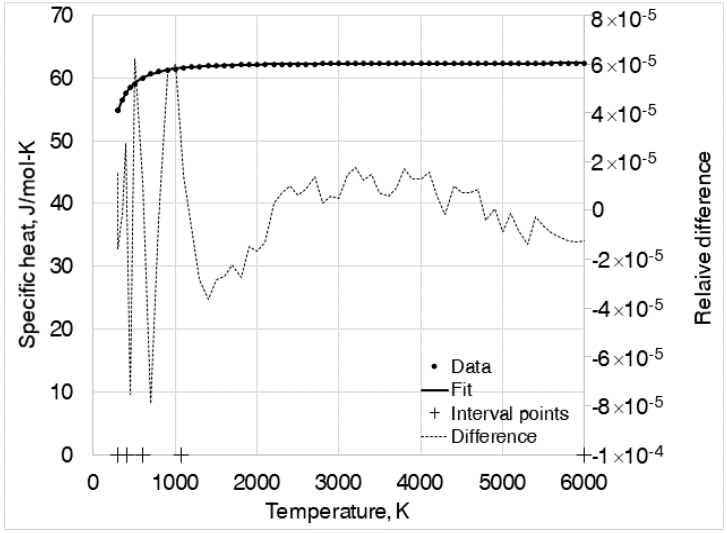


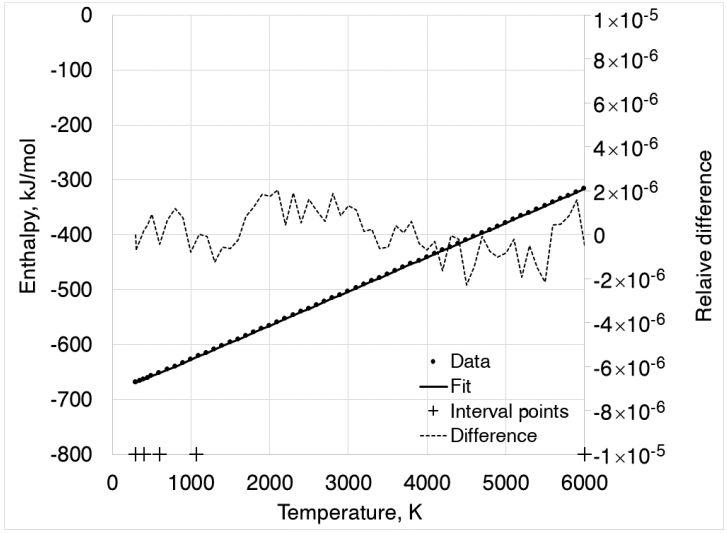


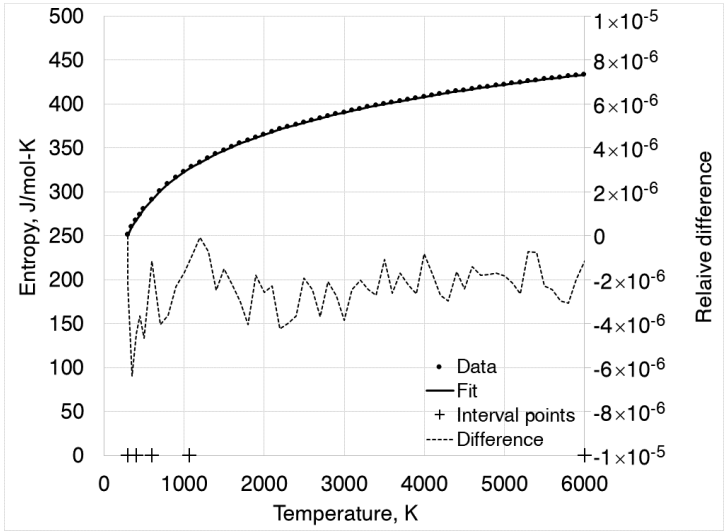


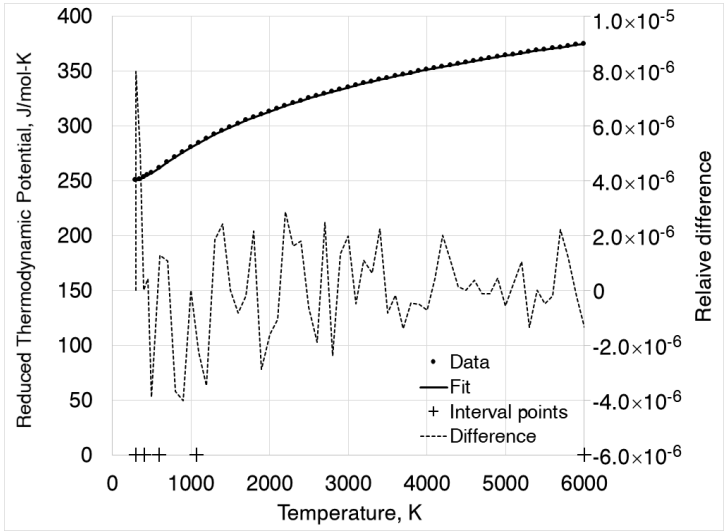
Sodium Fluoride, Ion (NaF_2^-) – F2Na1-(g)

MW = 60.9871 g/mol – $\Delta_f H^o$(298.15 K) = -669.44 kJ/mol – S^o(298.15 K) = 250.672 J/mol-K – p^o = 0.1 MPa



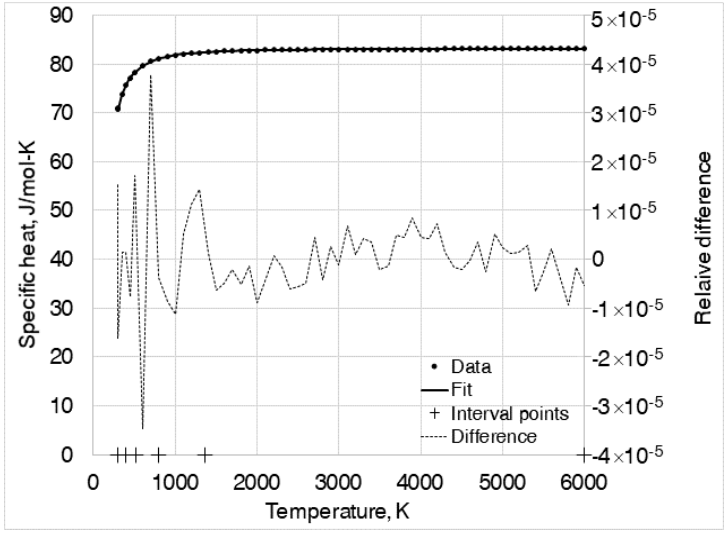


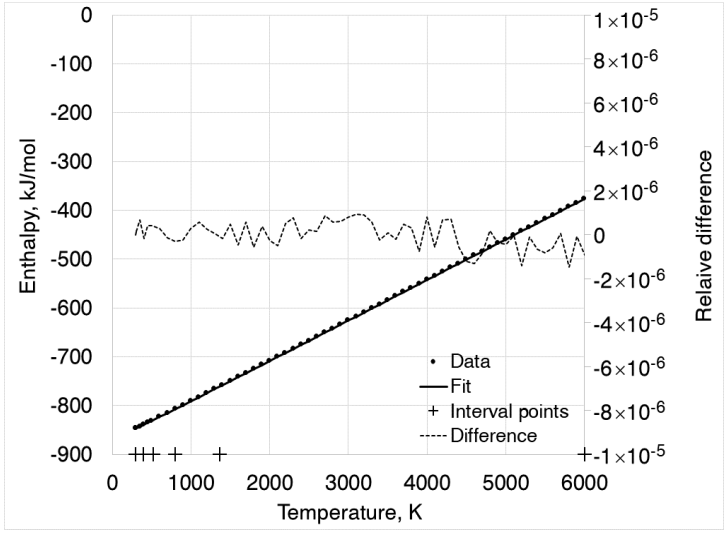


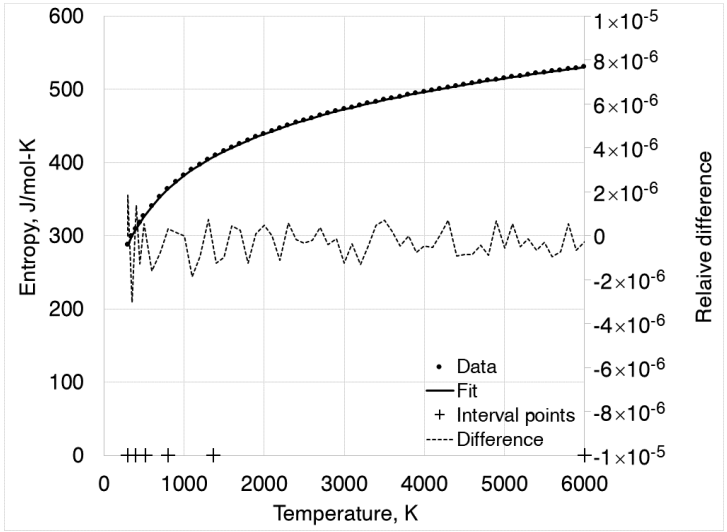


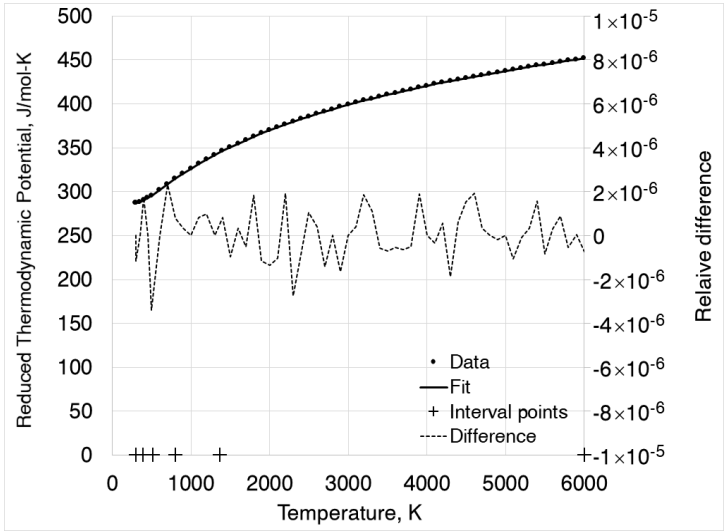
Sodium Fluoride ($(NaF)_2$) – F2Na2(g)

MW = 83.9763 g/mol – $\Delta_f H^o$(298.15 K) = -846.423 kJ/mol – S^o(298.15 K) = 287.366 J/mol-K – p^o = 0.1 MPa



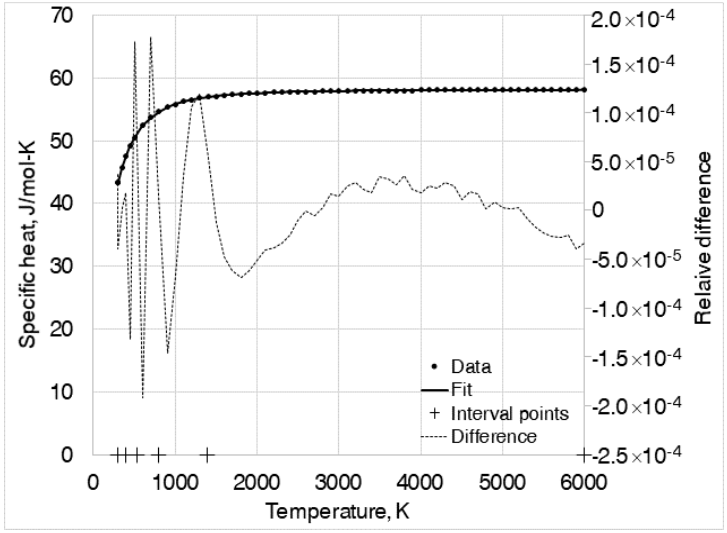


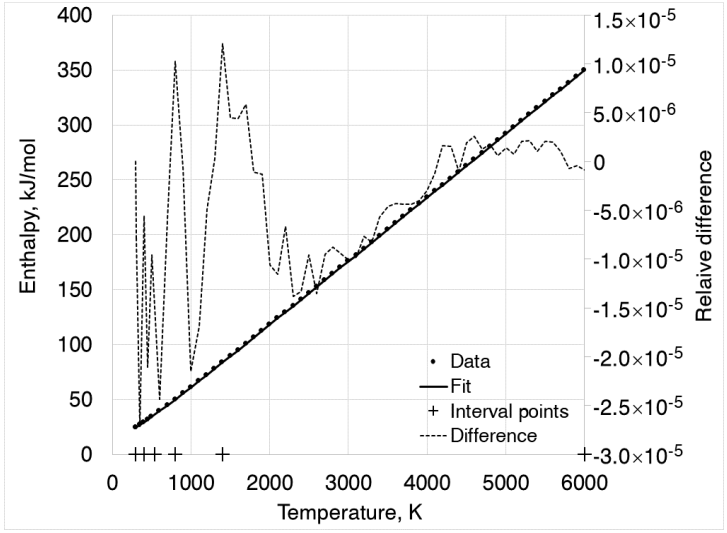


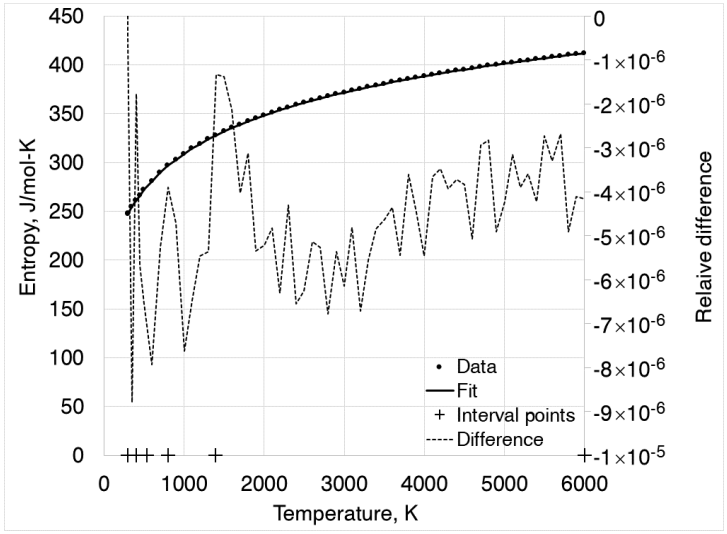


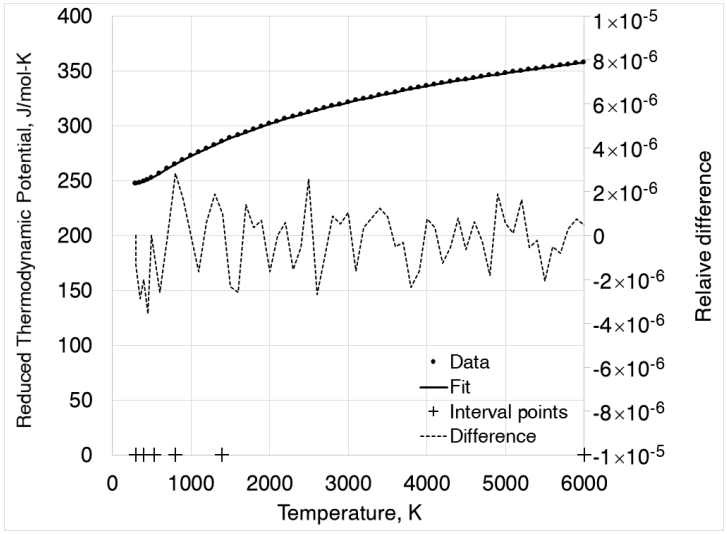
Oxygen Fluoride (OF_2) – F2O1(g)

MW = 53.9962 g/mol – $\Delta_f H^o$(298.15 K) = 24.518 kJ/mol – S^o(298.15 K) = 247.457 J/mol-K – p^o = 0.1 MPa



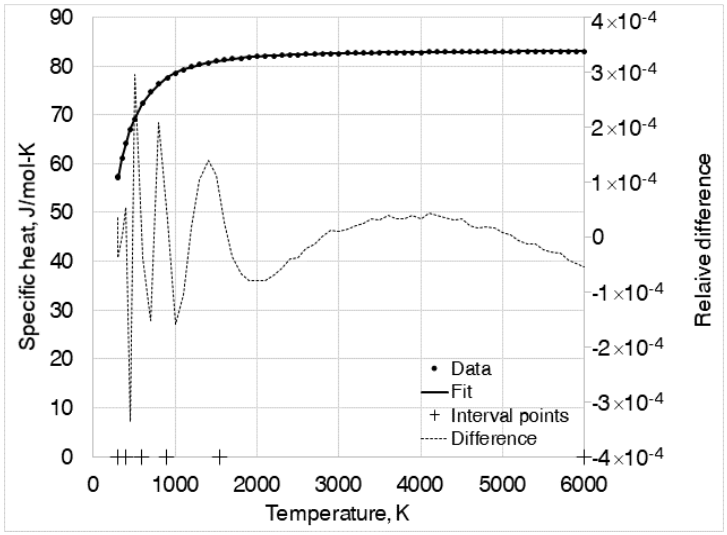


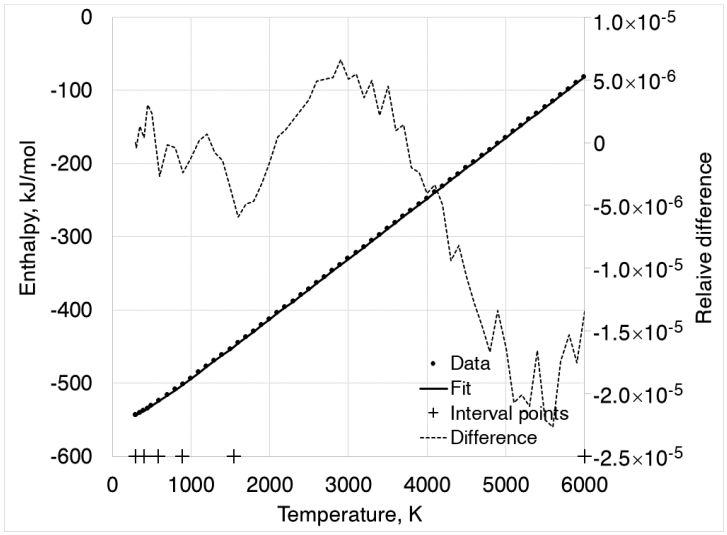


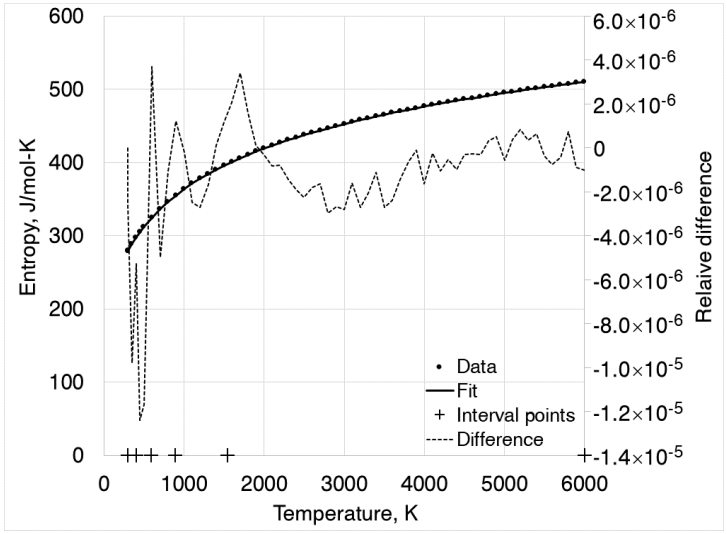


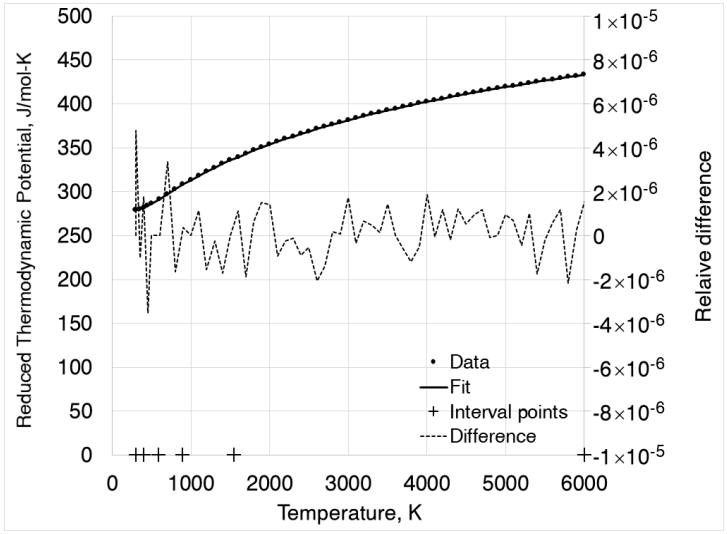
Thionyl Fluoride (OSF_2) – F2O1S1(g)

MW = 86.06 g/mol – $\Delta_f H^o$(298.15 K) = -543.92 kJ/mol – S^o(298.15 K) = 279.132 J/mol-K – p^o = 0.1 MPa



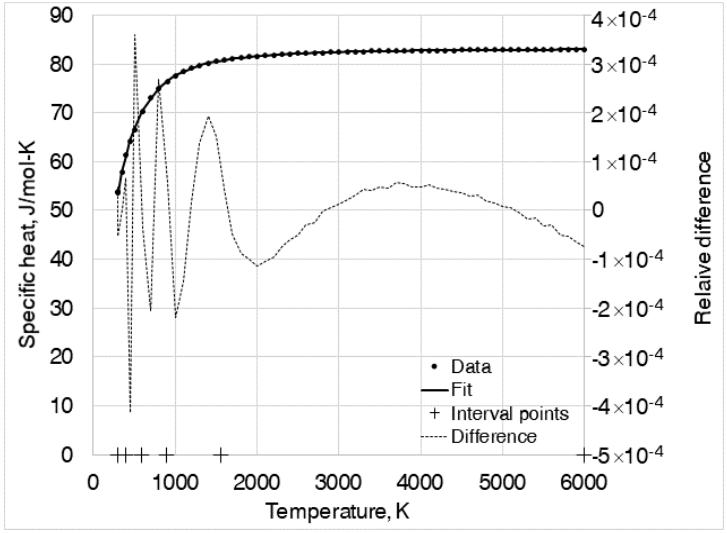


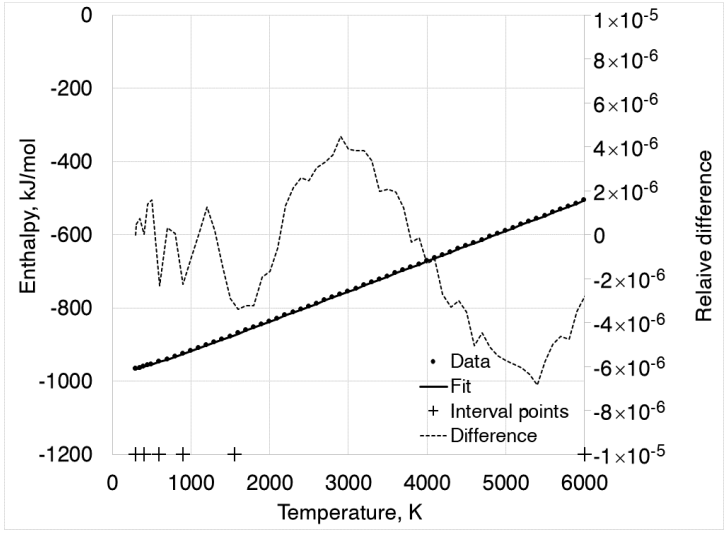


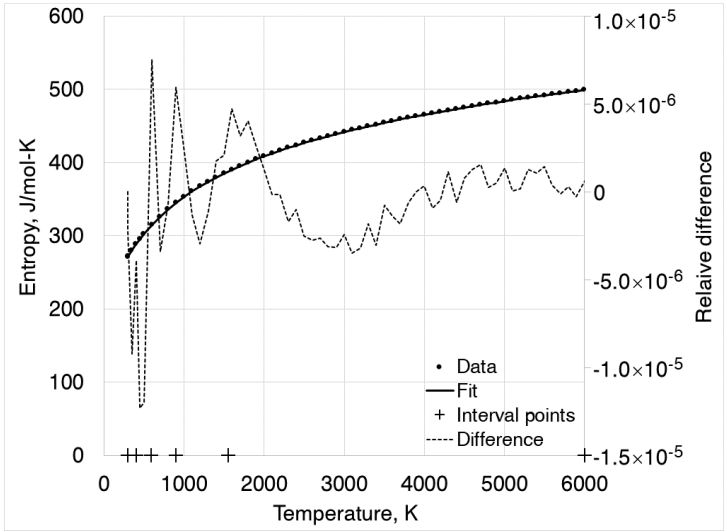


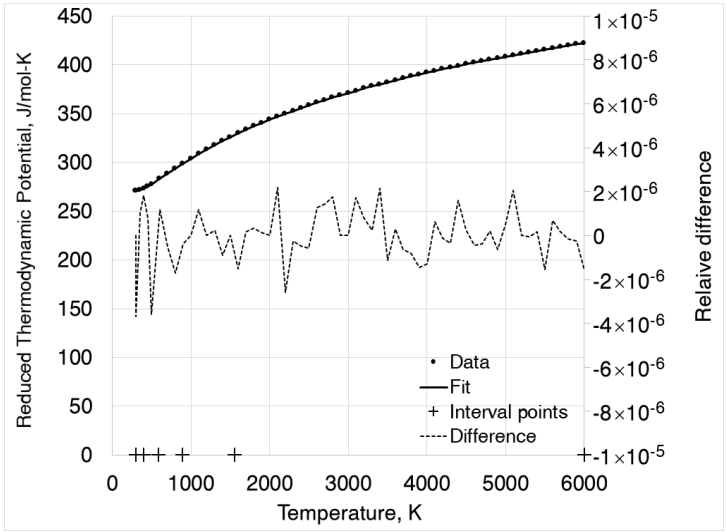
Difluorooxosilane ($OSiF_2$) – F2O1Si1(g)

MW = 82.0817 g/mol – $\Delta_f H^o$(298.15 K) = -966.504 kJ/mol – S^o(298.15 K) = 271.247 J/mol-K – p^o = 0.1 MPa



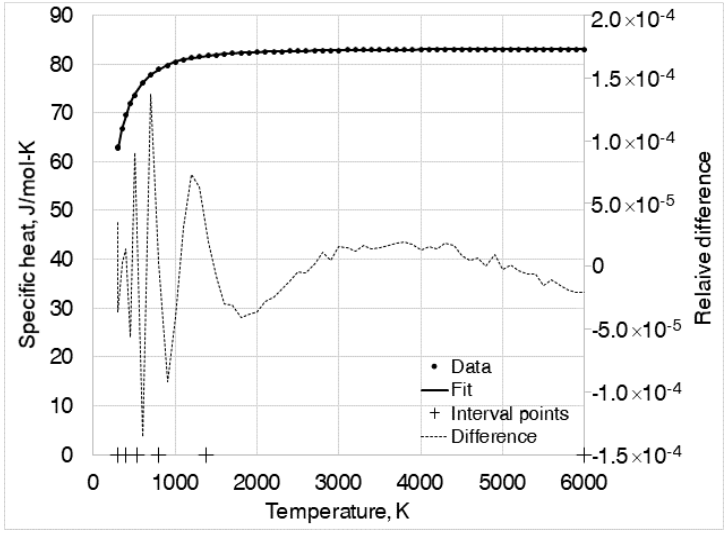


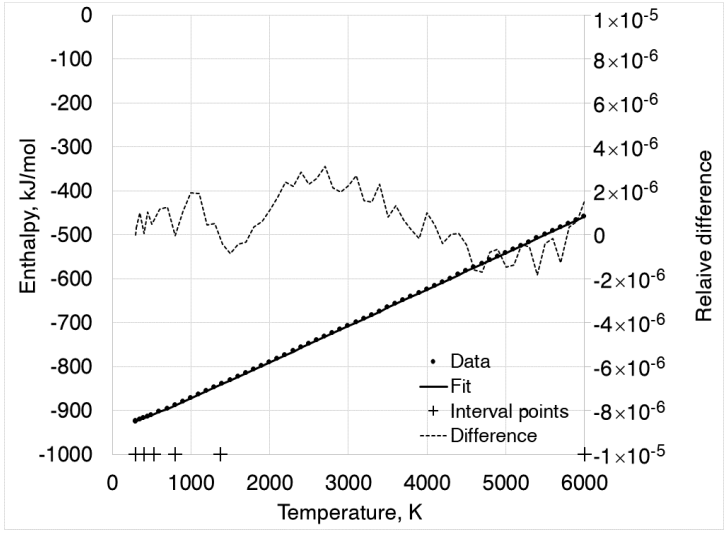


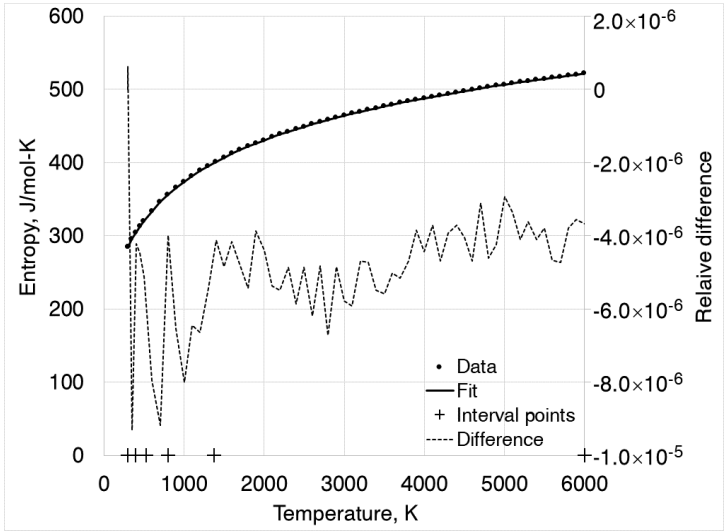


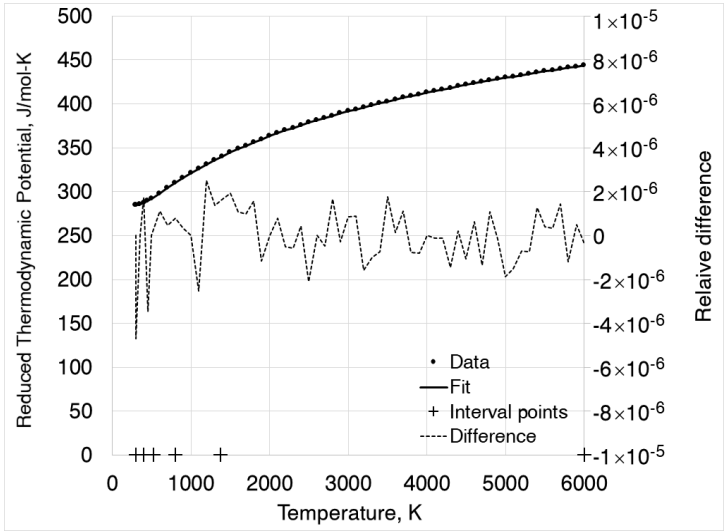
Titanium Fluoride Oxide ($OTiF_2$) – F2O1Ti1(g)

MW = 101.88 g/mol – $\Delta_f H^o$(298.15 K) = -924.873 kJ/mol – S^o(298.15 K) = 284.678 J/mol-K – p^o = 0.1 MPa



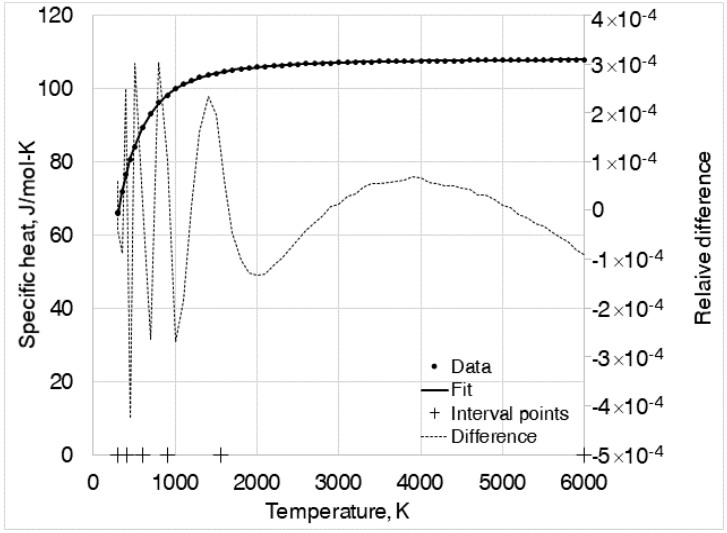


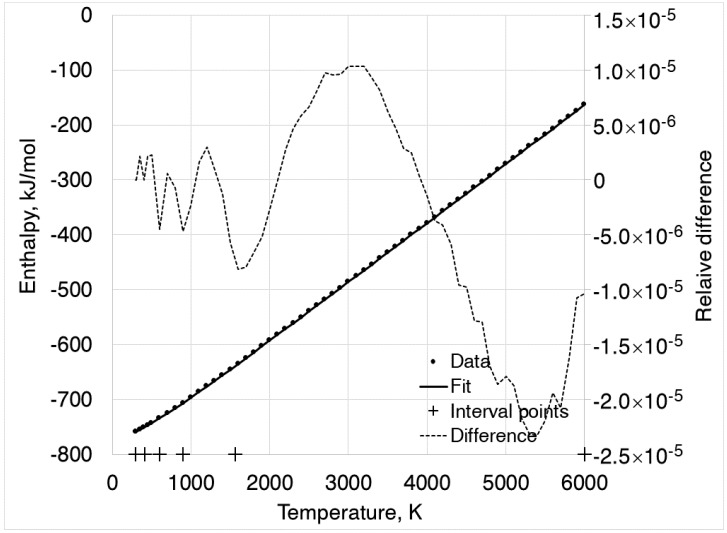


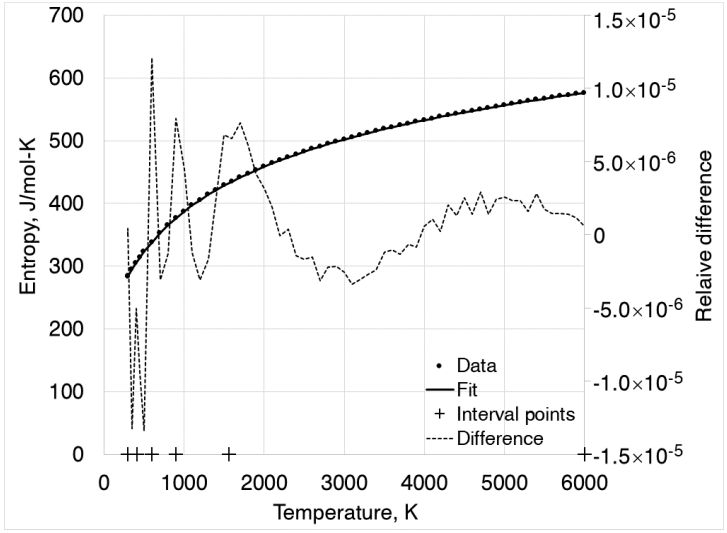


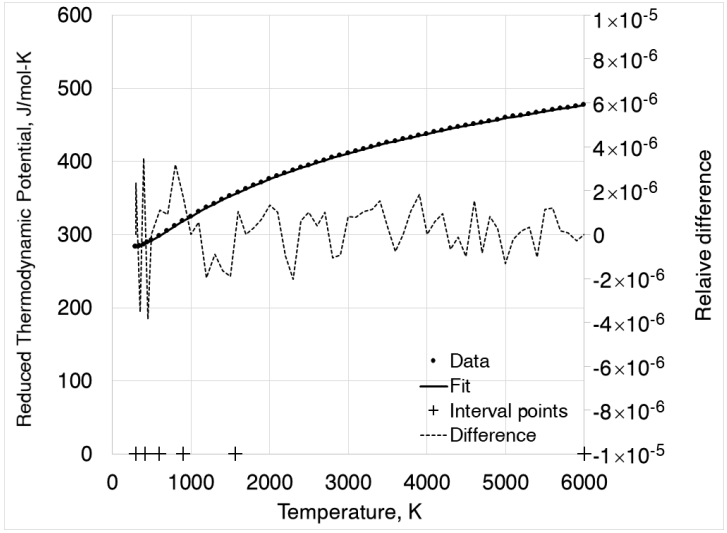
Sulfuryl Fluoride (SO_2F_2) – F2O2S1(g)

MW = 102.06 g/mol – $\Delta_f H^o$(298.15 K) = -758.559 kJ/mol – S^o(298.15 K) = 283.623 J/mol-K – p^o = 0.1 MPa



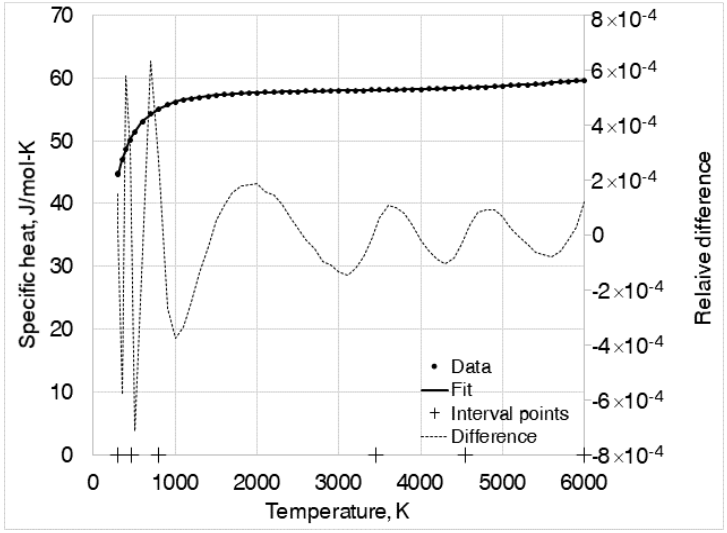


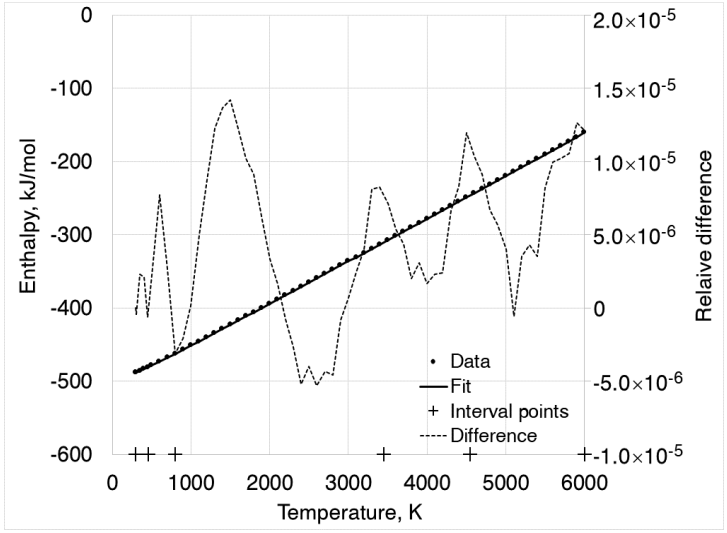


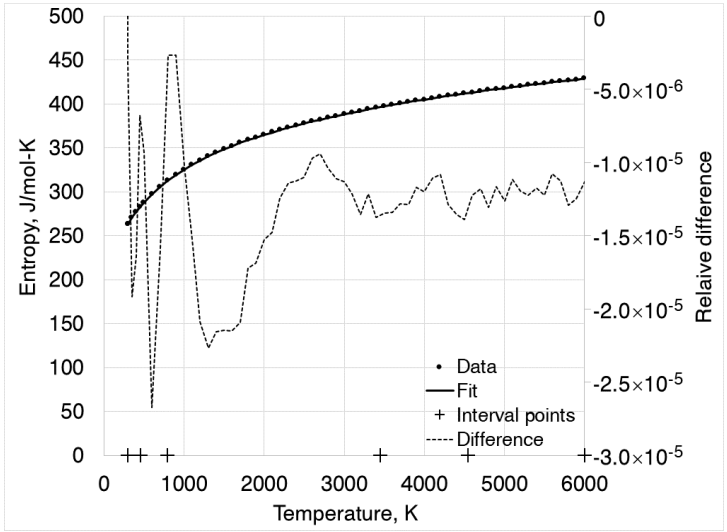


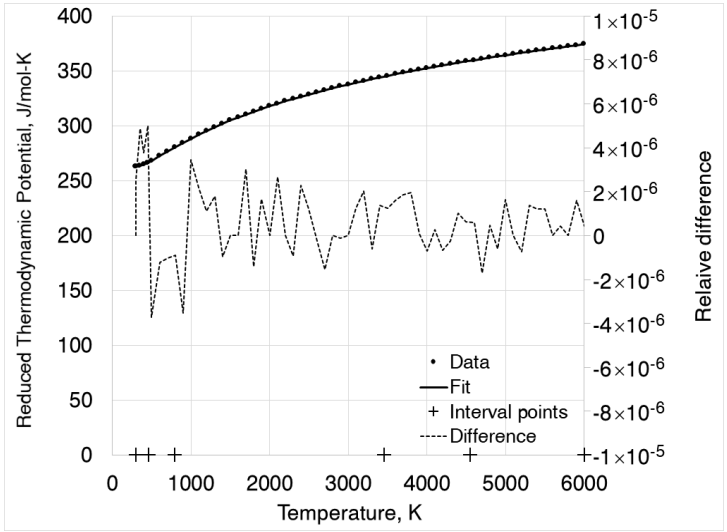
Phosphorus Fluoride (PF_2) – F2P1(g)

MW = 68.9706 g/mol – $\Delta_f H^o$(298.15 K) = -488.256 kJ/mol – S^o(298.15 K) = 262.957 J/mol-K – p^o = 0.1 MPa



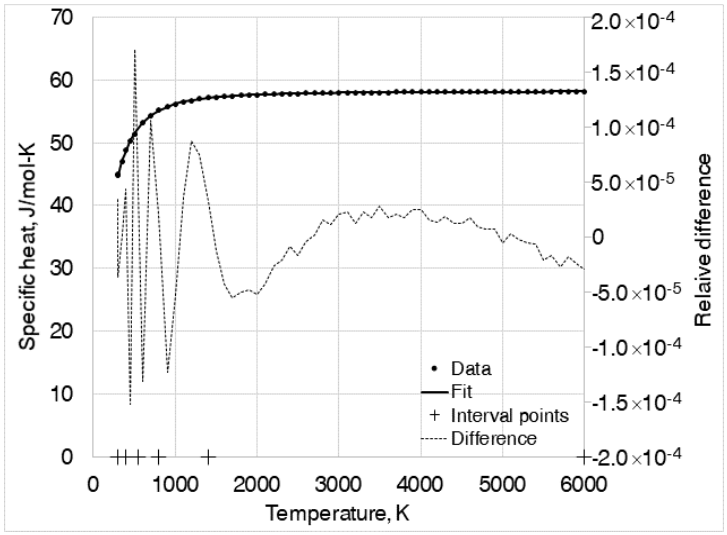


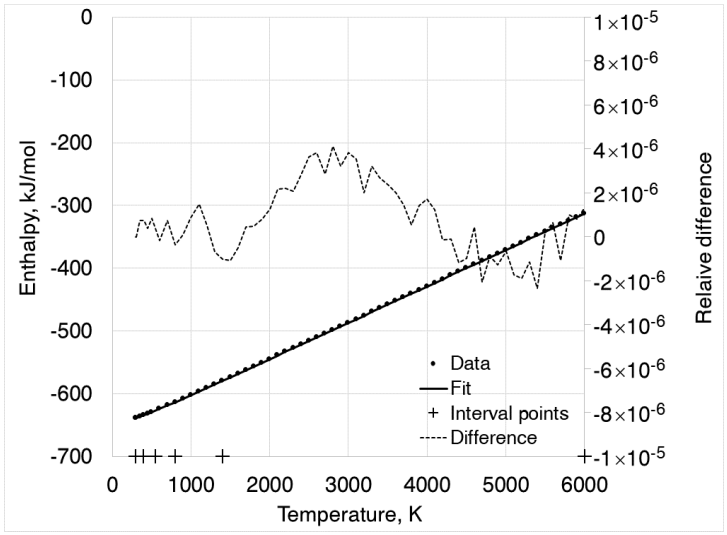


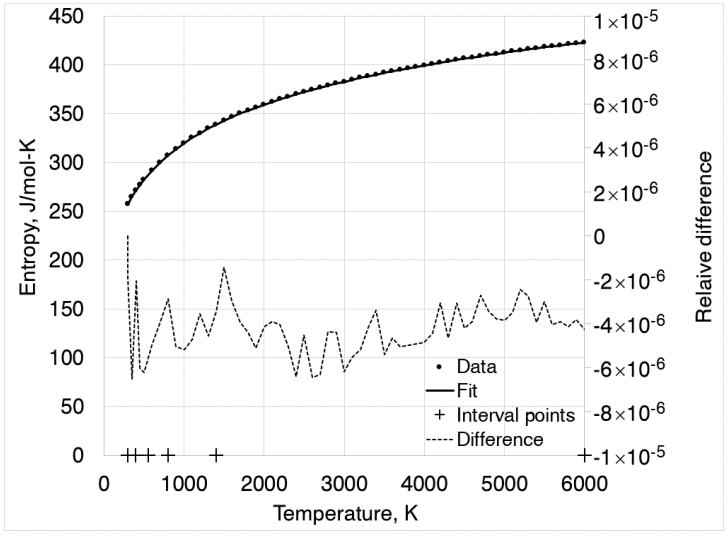


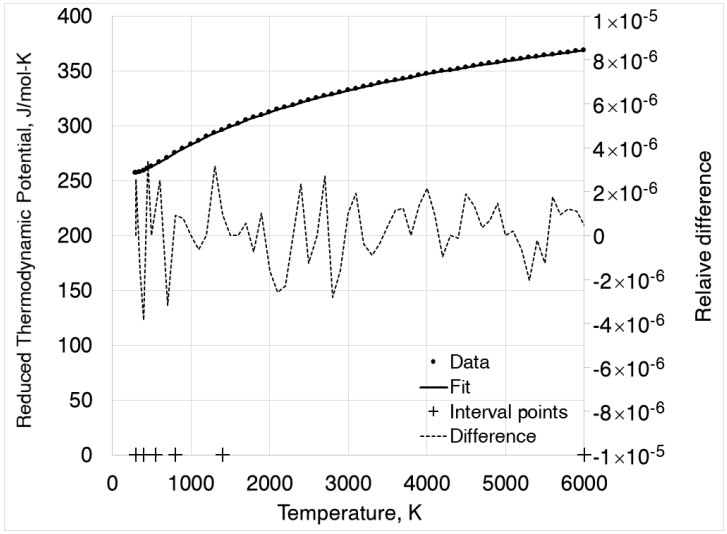
Phosphorus Fluoride, Ion (PF_2^-) – F2P1-(g)

MW = 68.9711 g/mol – $\Delta_f H^o$(298.15 K) = -639.177 kJ/mol – S^o(298.15 K) = 257.364 J/mol-K – p^o = 0.1 MPa



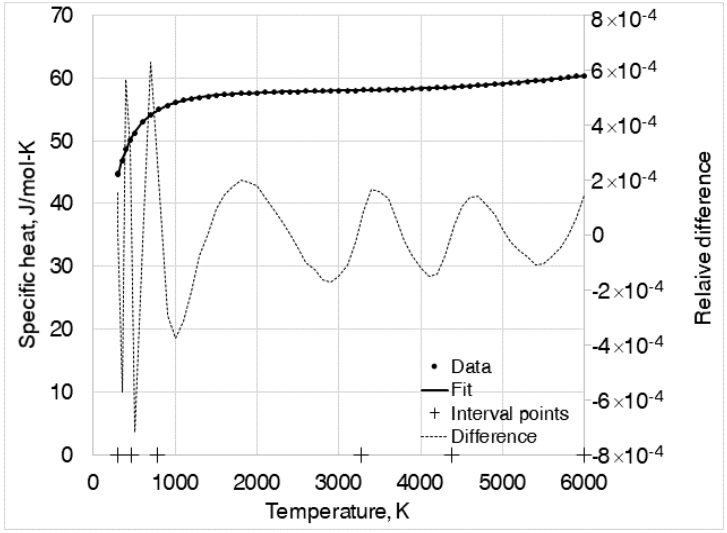


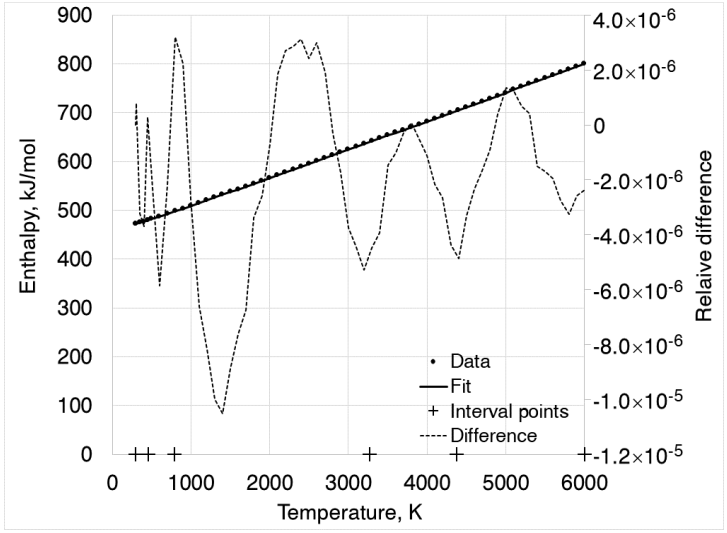


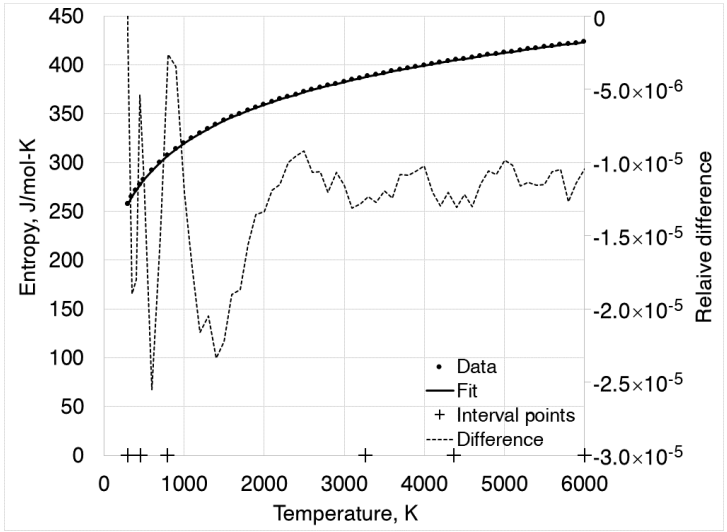


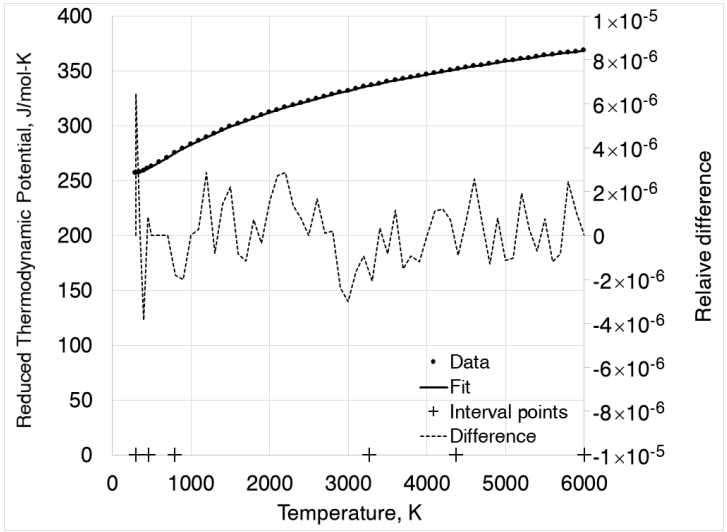
Phosphorus Fluoride, Ion (PF_2^+) – F2P1+(g)

MW = 68.97 g/mol – $\Delta_f H^o$(298.15 K) = 472.273 kJ/mol – S^o(298.15 K) = 257.358 J/mol-K – p^o = 0.1 MPa



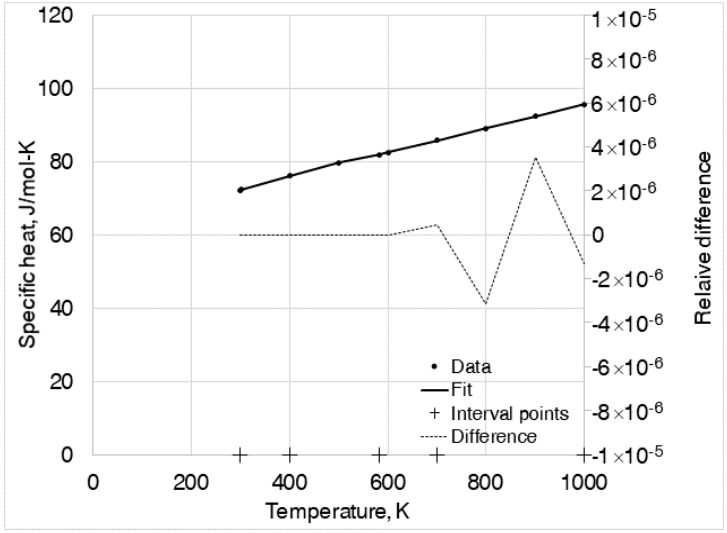


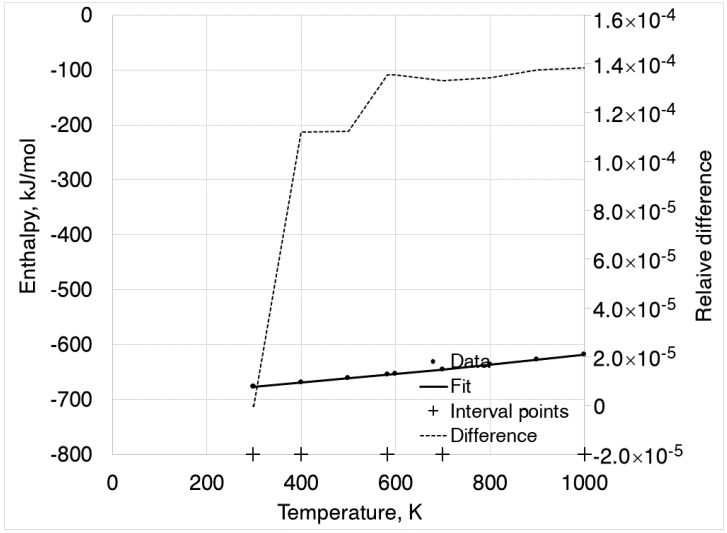


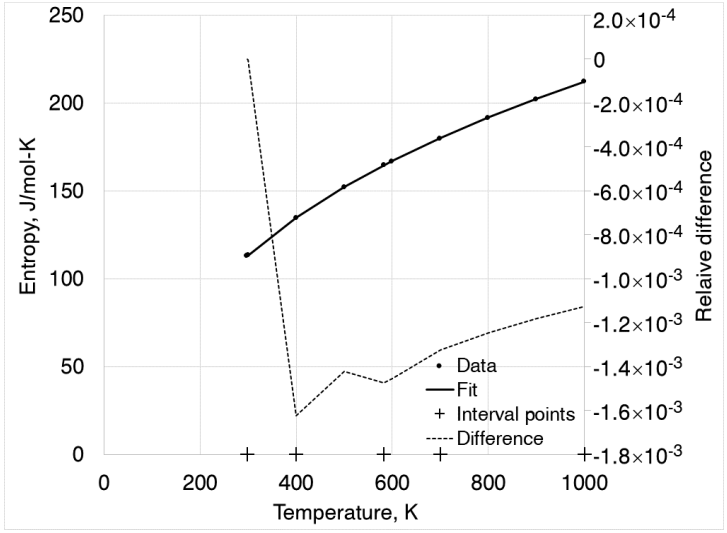


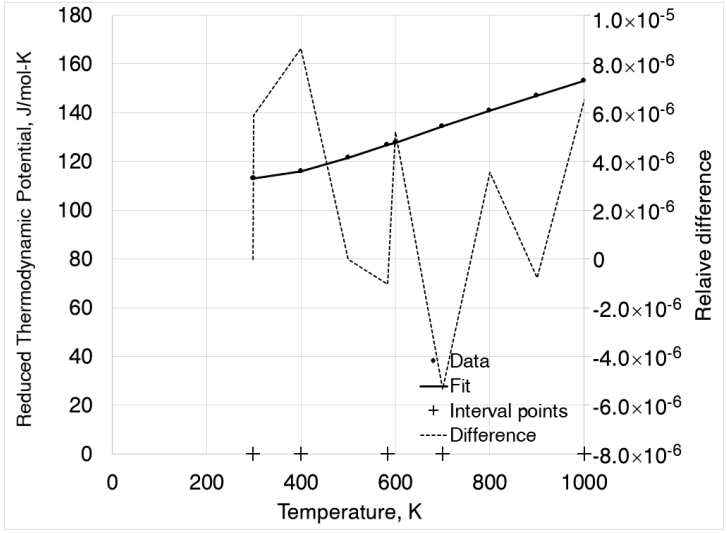
Lead Fluoride, Alpha (PbF_2) – F2Pb1(cr)_01

MW = 245.2 g/mol – $\Delta_f H^o$(298.15 K) = -676.971 kJ/mol – S^o(298.15 K) = 112.968 J/mol-K – p^o = 0.1 MPa



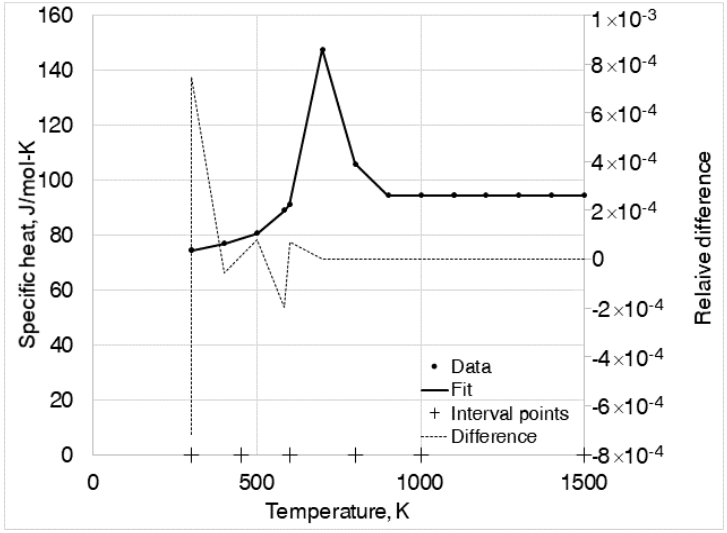


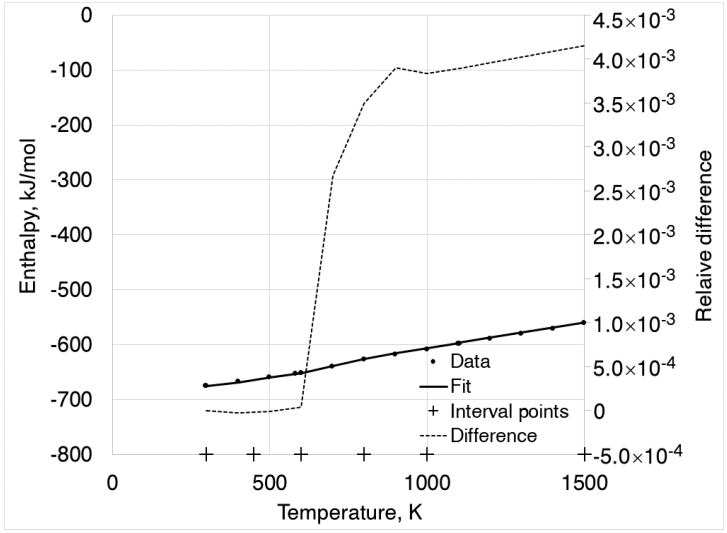


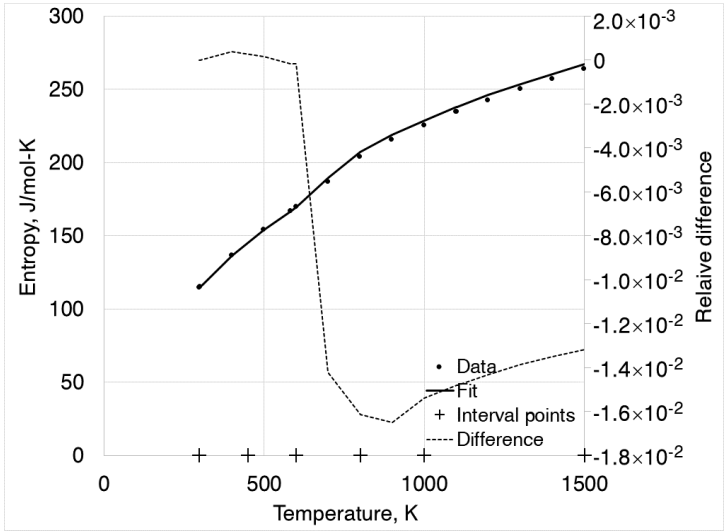


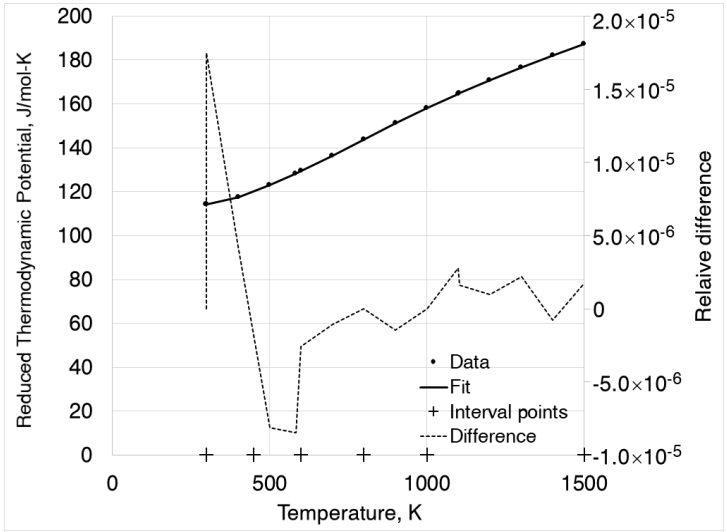
Lead Fluoride, Beta (PbF_2) – F2Pb1(cr)_02

MW = 245.2 g/mol – $\Delta_f H^o$(298.15 K) = -675.99 kJ/mol – S^o(298.15 K) = 114.398 J/mol-K – p^o = 0.1 MPa



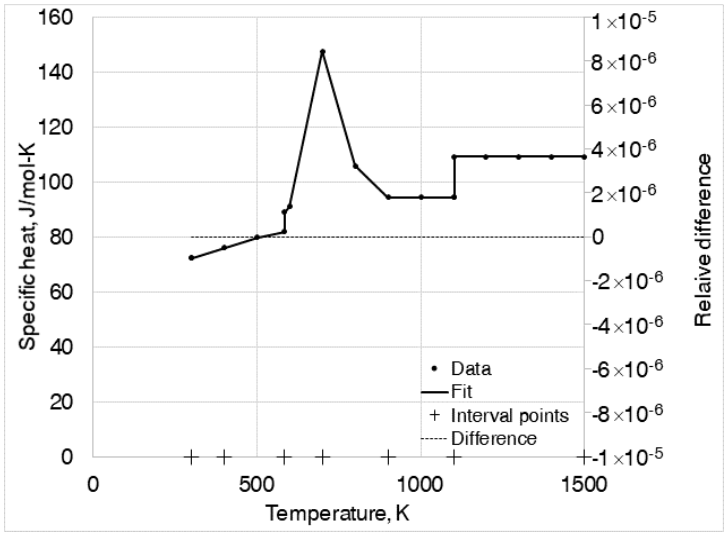


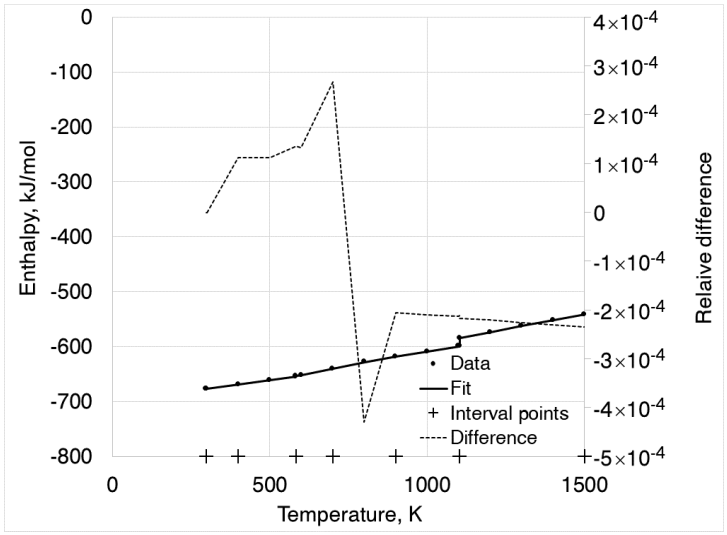


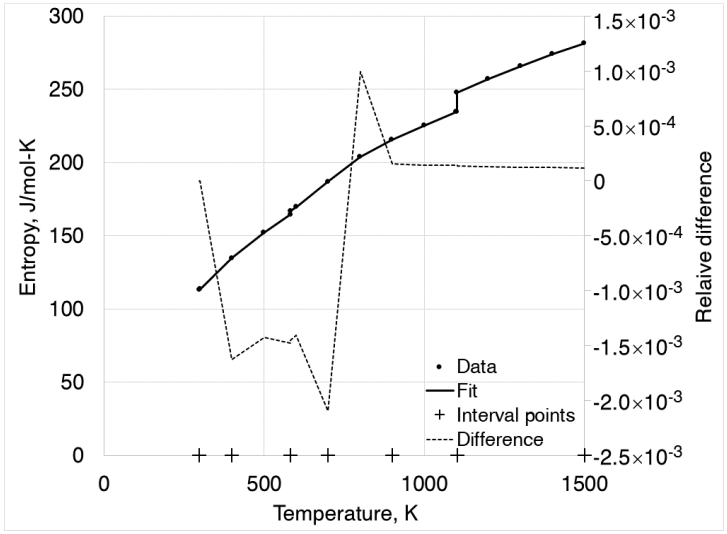


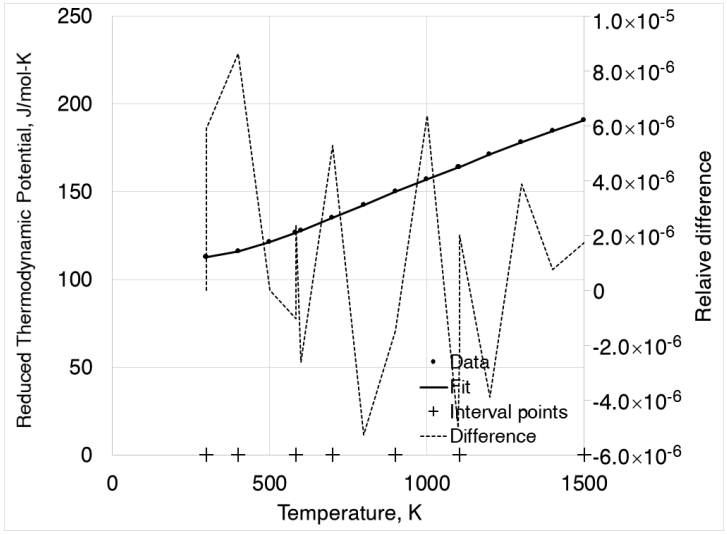
Lead Fluoride (PbF_2) – F2Pb1(cr,l)

MW = 245.2 g/mol – $\Delta_f H^o$(298.15 K) = -676.971 kJ/mol – S^o(298.15 K) = 112.968 J/mol-K – p^o = 0.1 MPa



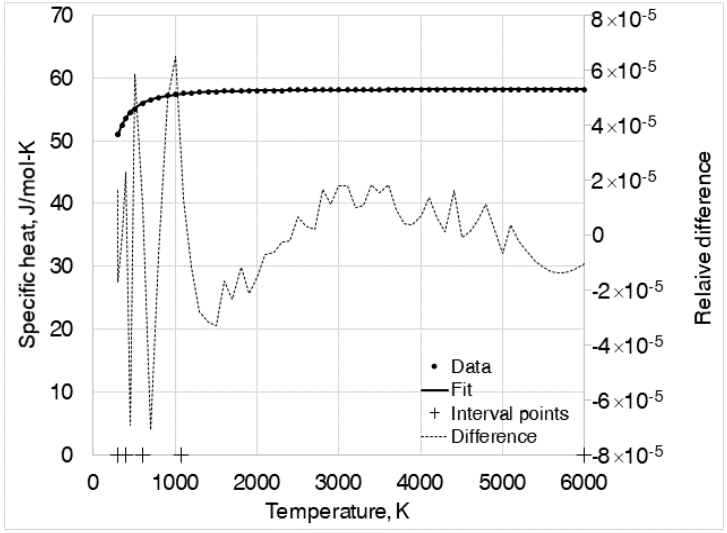


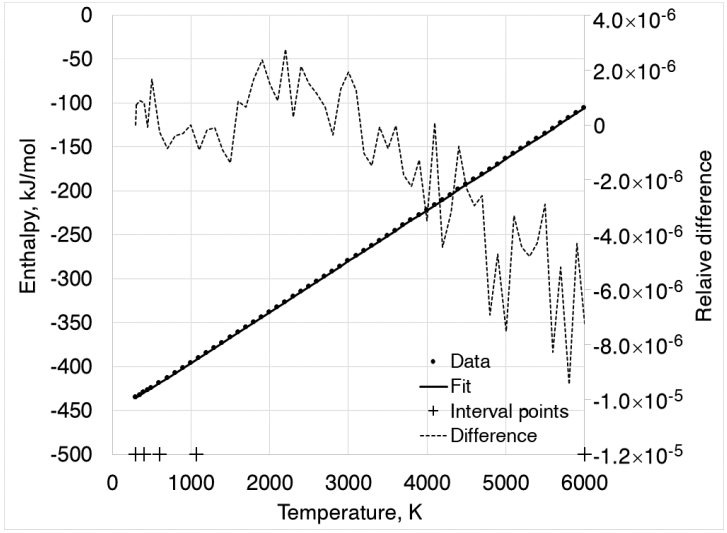


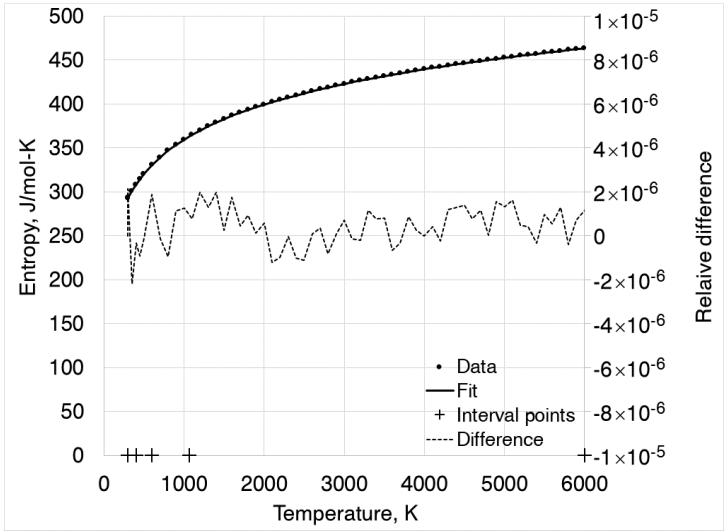


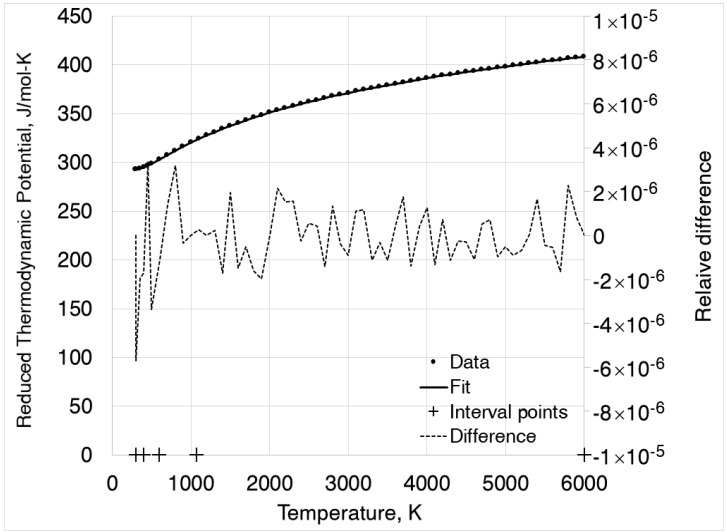
Lead Fluoride (PbF_2) – F2Pb1(g)

MW = 245.2 g/mol – $\Delta_f H^o$(298.15 K) = -435.136 kJ/mol – S^o(298.15 K) = 292.692 J/mol-K – p^o = 0.1 MPa



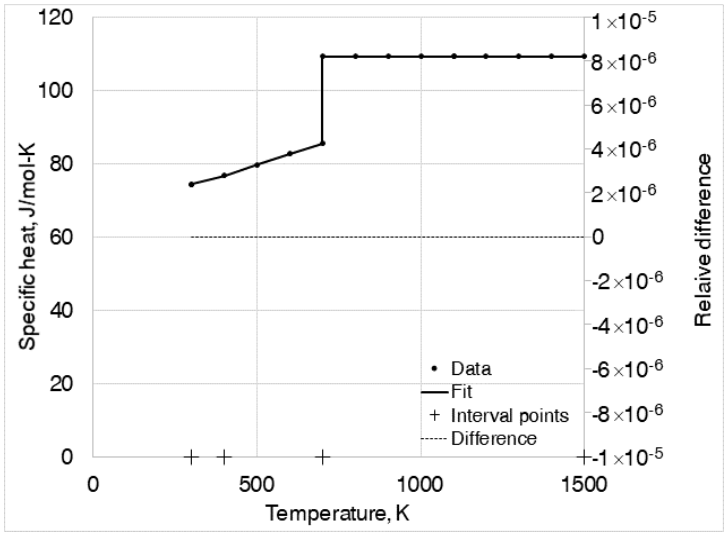


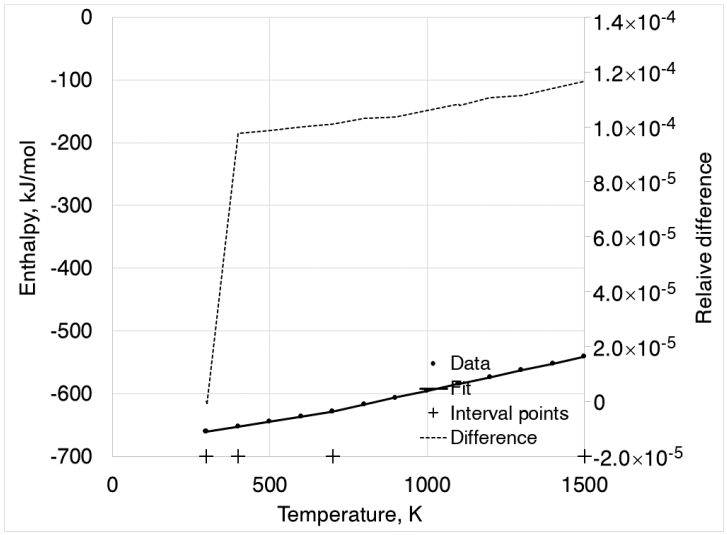


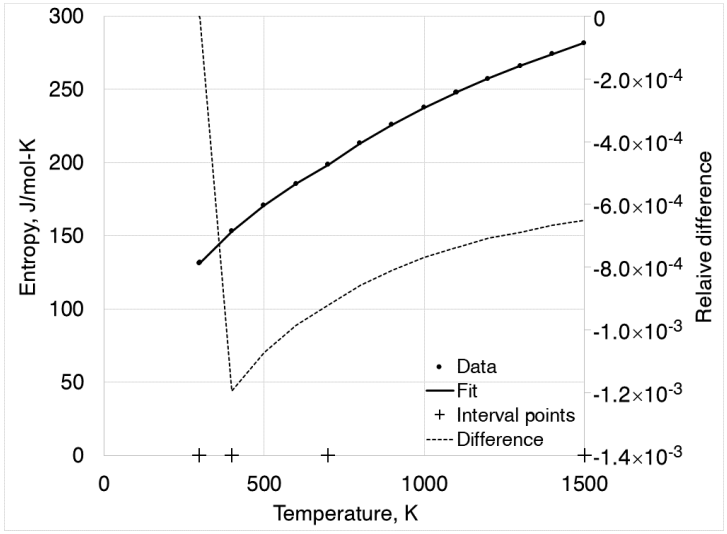


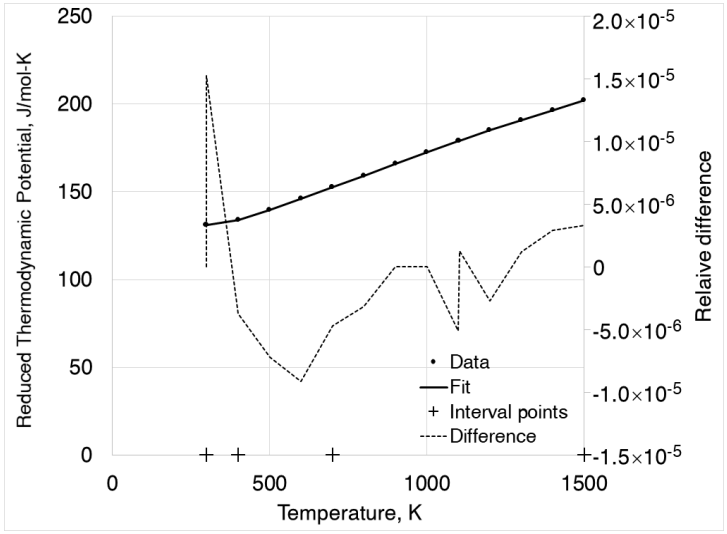
Lead Fluoride (PbF_2) – F2Pb1(l)

MW = 245.2 g/mol – $\Delta_f H^o$(298.15 K) = -660.32 kJ/mol – S^o(298.15 K) = 130.927 J/mol-K – p^o = 0.1 MPa



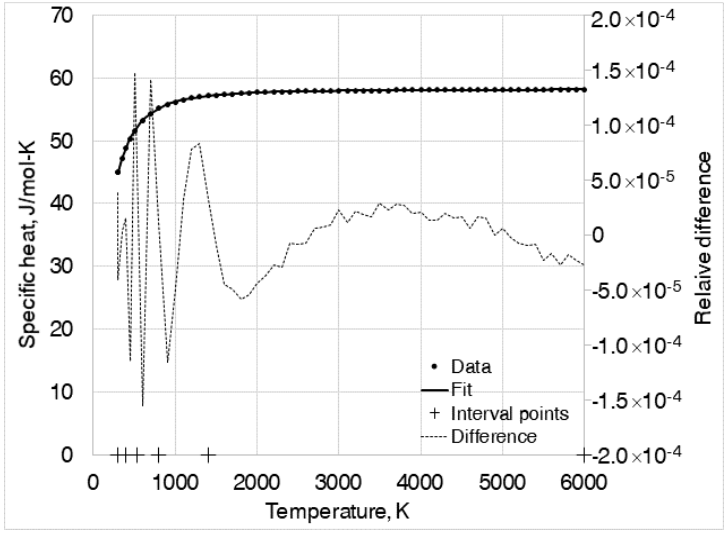


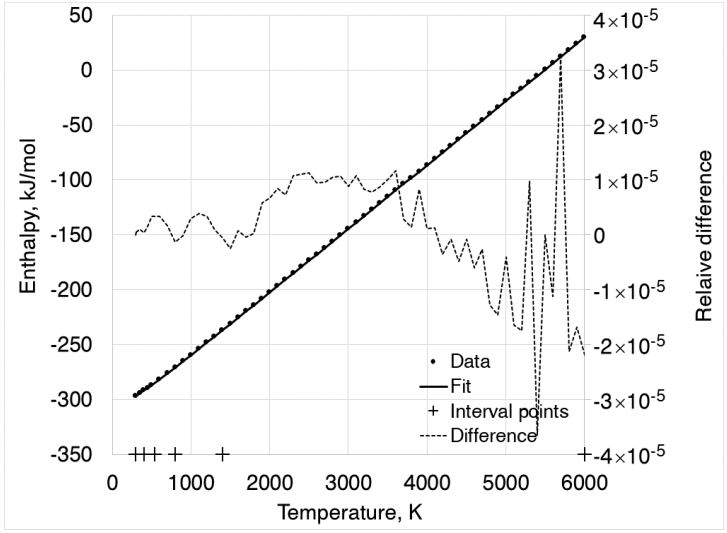


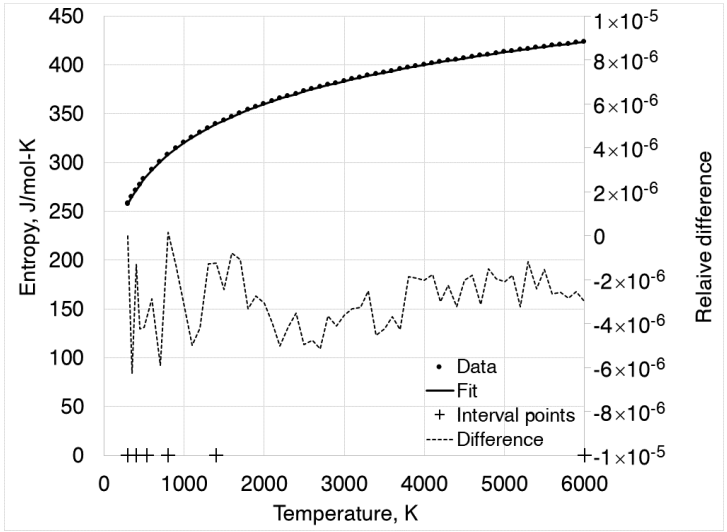


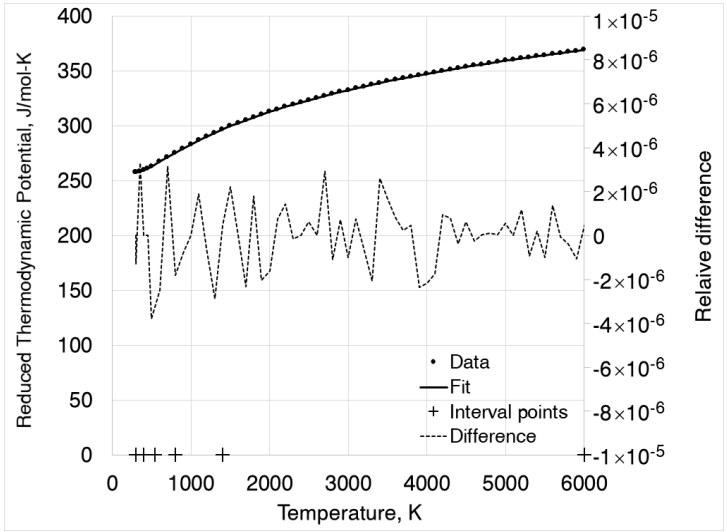
Sulfur Fluoride (SF_2) – F2S1(g)

MW = 70.06 g/mol – $\Delta_f H^o$(298.15 K) = -296.646 kJ/mol – S^o(298.15 K) = 257.703 J/mol-K – p^o = 0.1 MPa



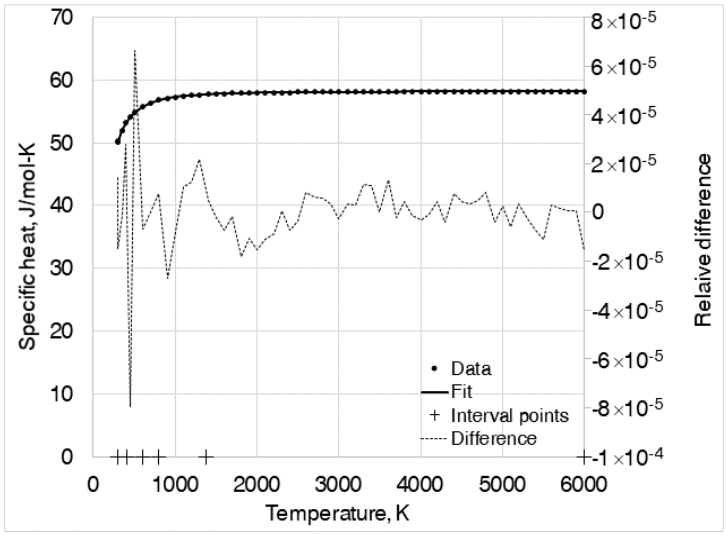


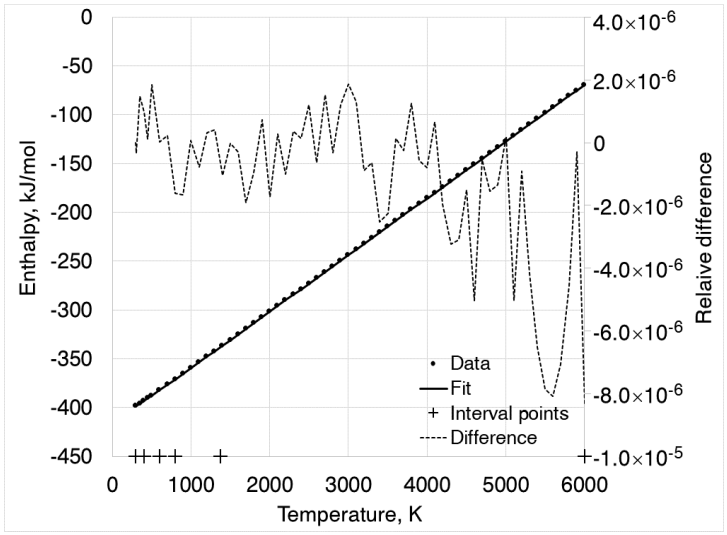


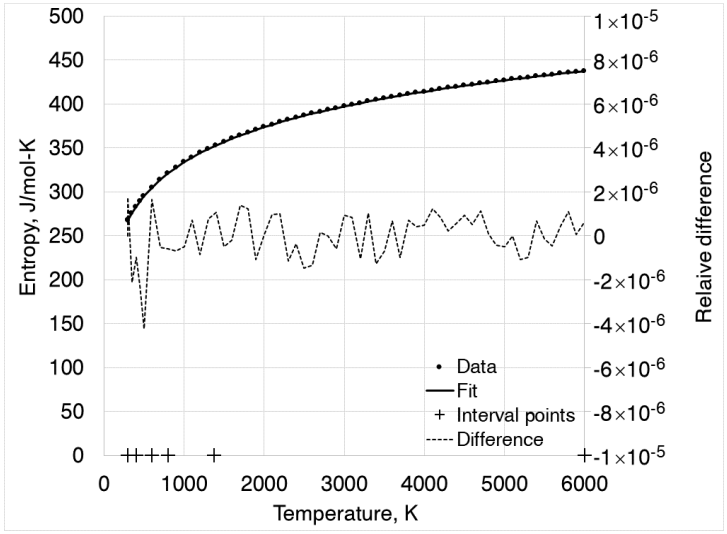


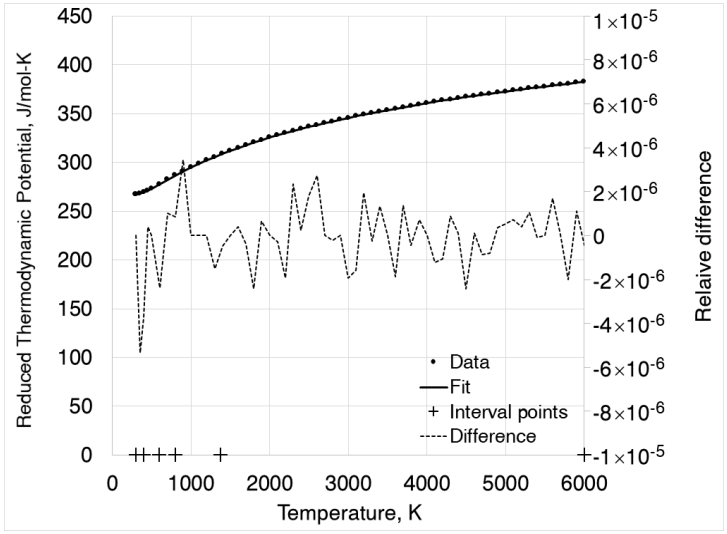
Sulfur Fluoride, Ion (SF_2^-) – F2S1-(g)

MW = 70.06 g/mol – $\Delta_f H^o$(298.15 K) = -398.426 kJ/mol – S^o(298.15 K) = 267.445 J/mol-K – p^o = 0.1 MPa



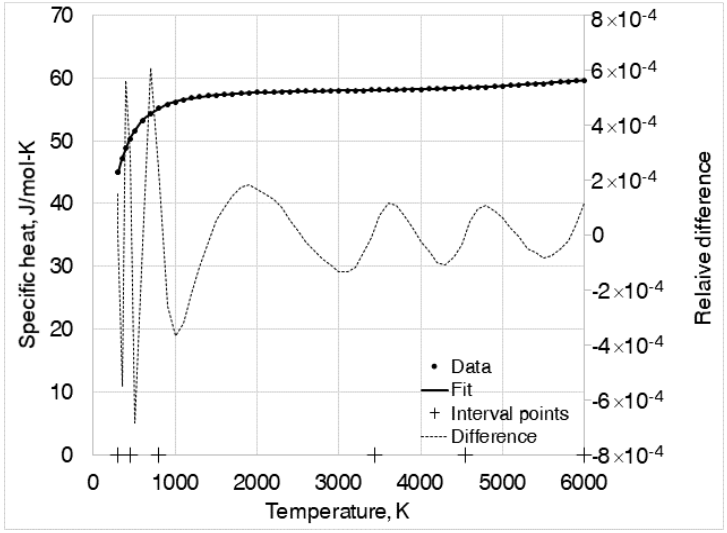


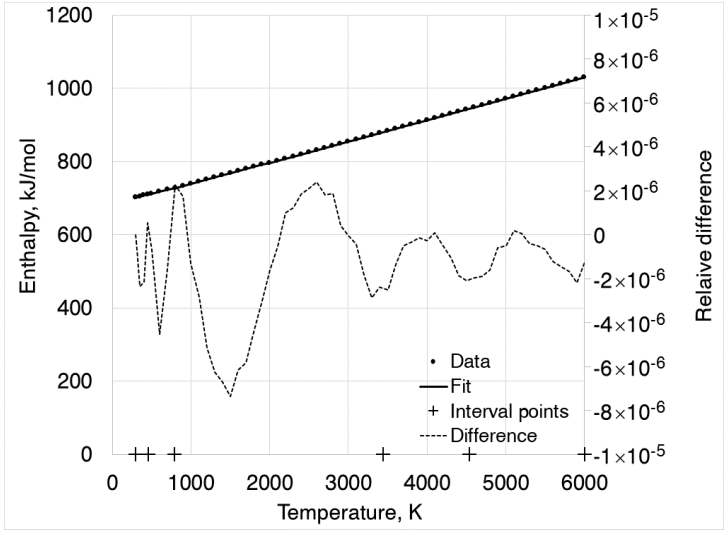


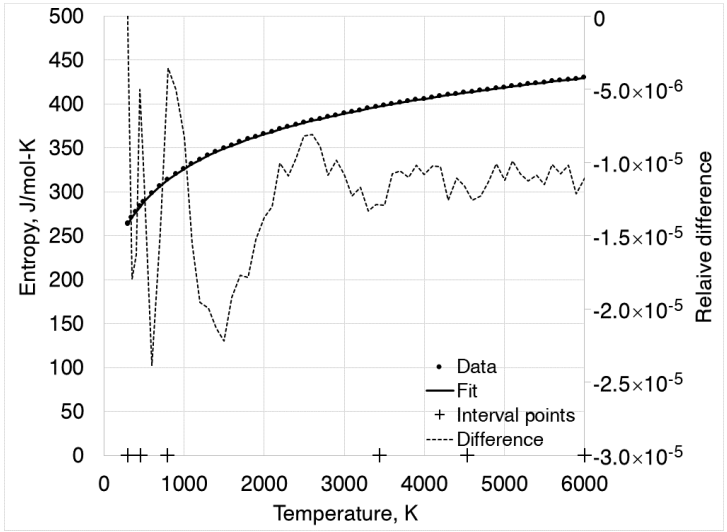


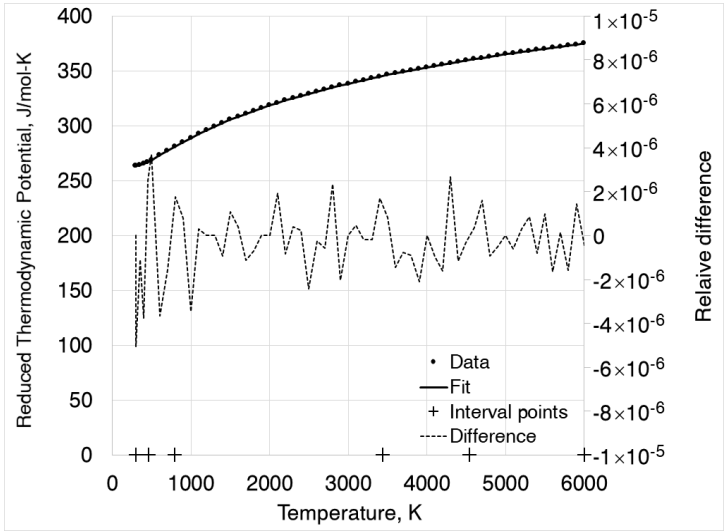
Sulfur Fluoride, Ion (SF_2^+) – F2S1+(g)

MW = 70.06 g/mol – $\Delta_f H^o$(298.15 K) = 702.389 kJ/mol – S^o(298.15 K) = 263.525 J/mol-K – p^o = 0.1 MPa



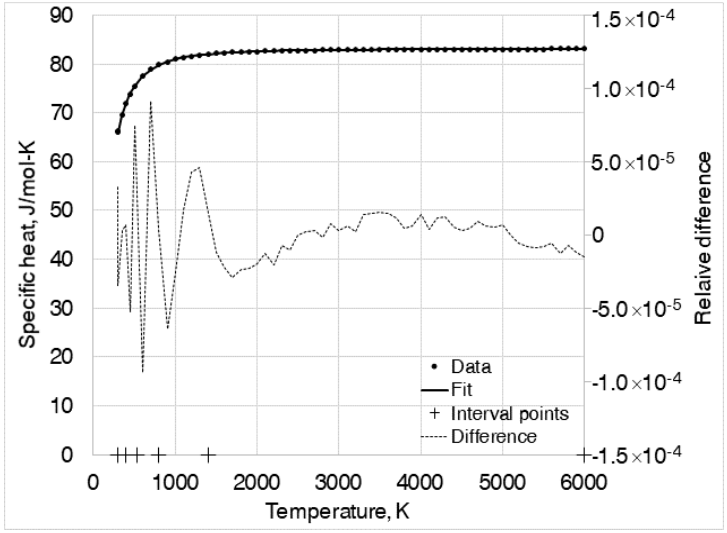


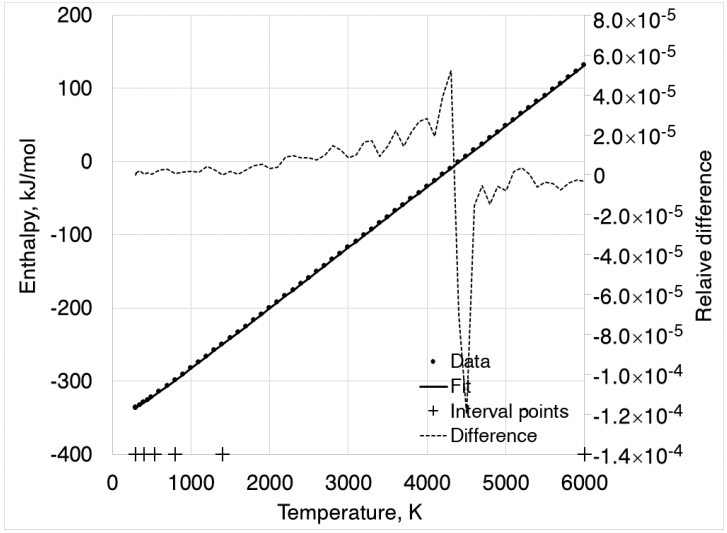


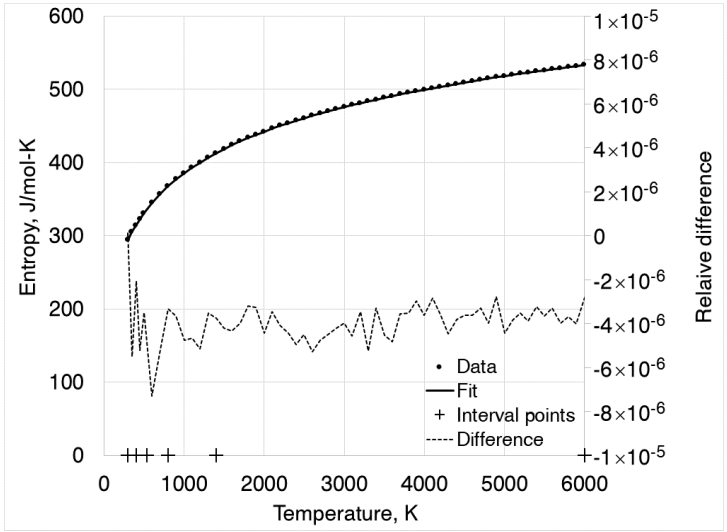


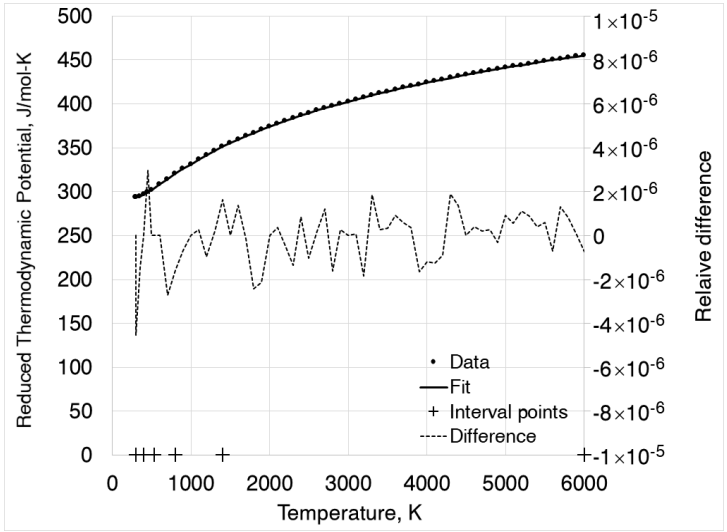
Difluorodisulfane (FSSF) – F2S2(g)_01

MW = 102.12 g/mol – $\Delta_f H^o$(298.15 K) = -336.435 kJ/mol – S^o(298.15 K) = 294.082 J/mol-K – p^o = 0.1 MPa



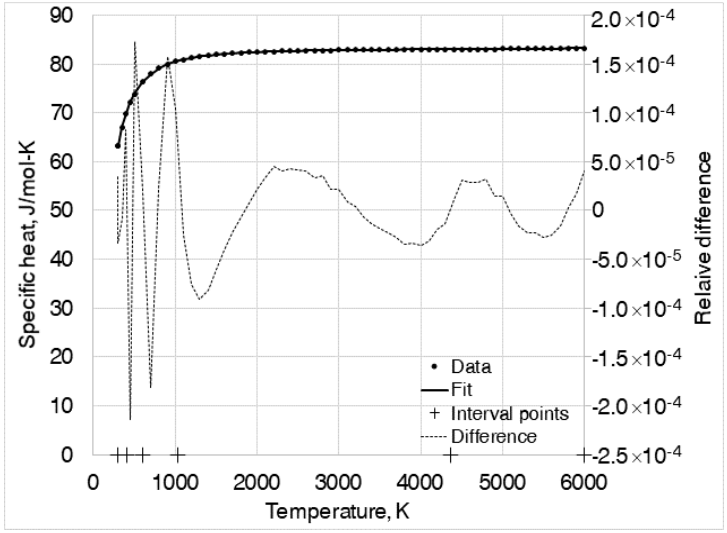


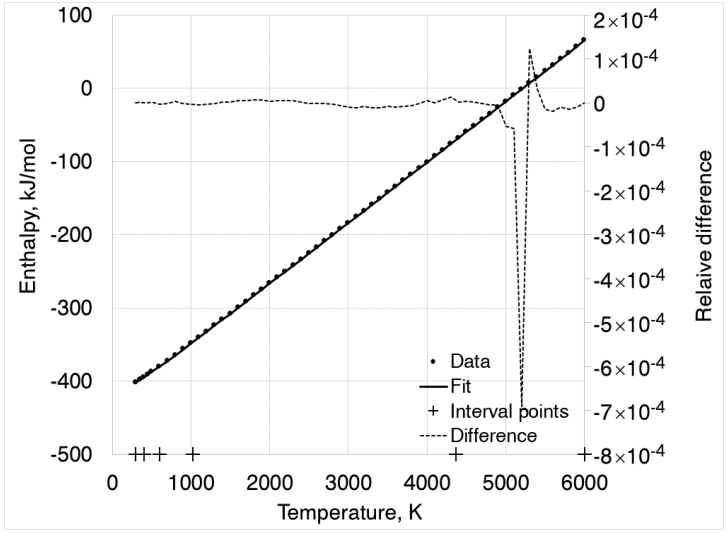


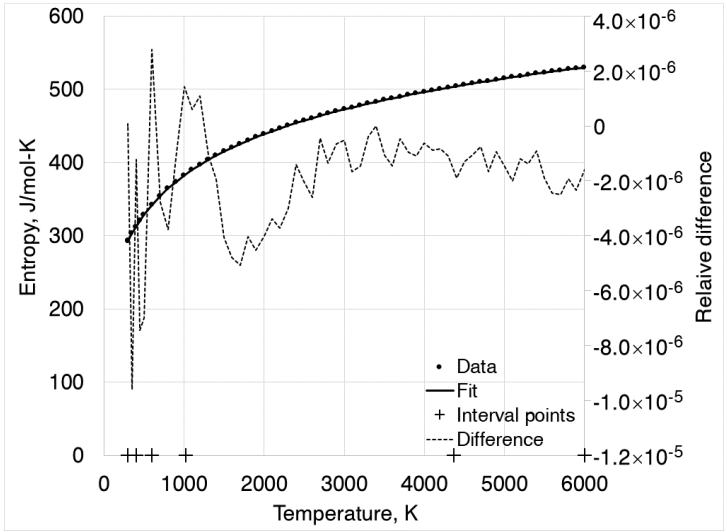


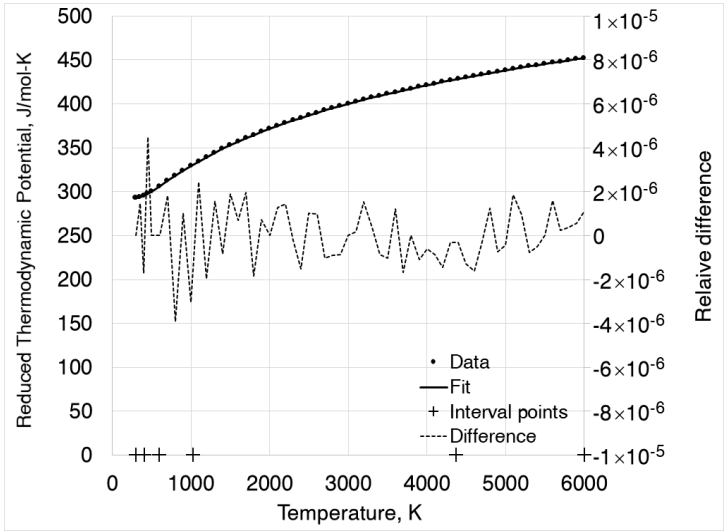
Thiothionyl Fluoride (SSF_2) – F2S2(g)_02

MW = 102.12 g/mol – $\Delta_f H^o$(298.15 K) = -401.413 kJ/mol – S^o(298.15 K) = 292.827 J/mol-K – p^o = 0.1 MPa



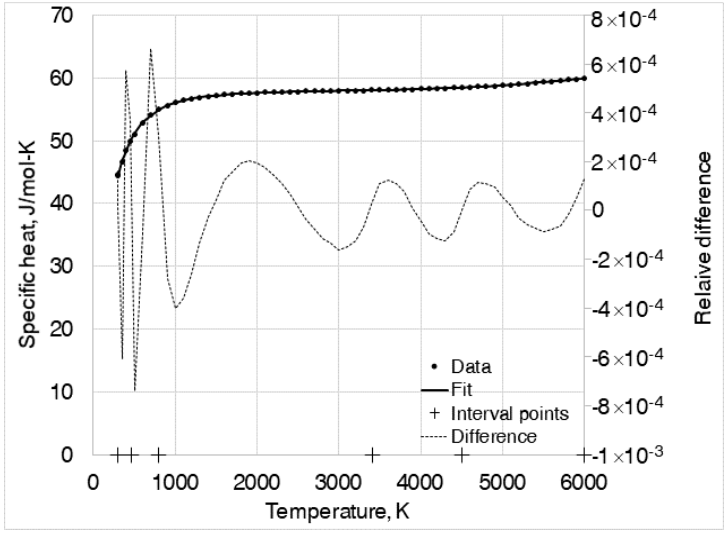


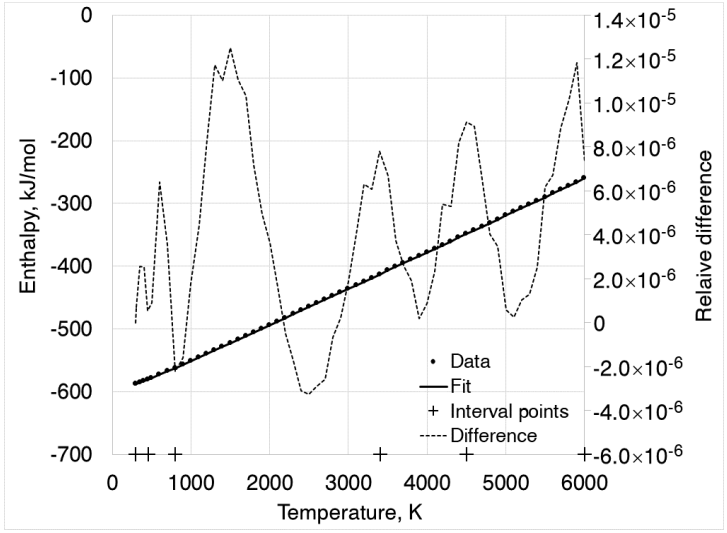


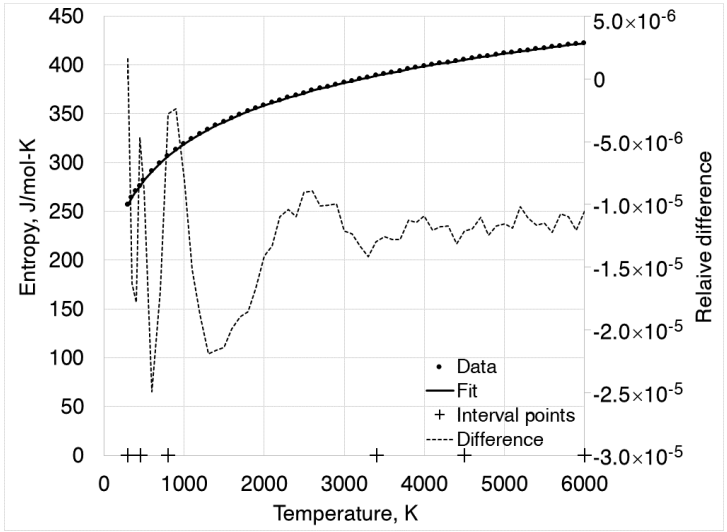


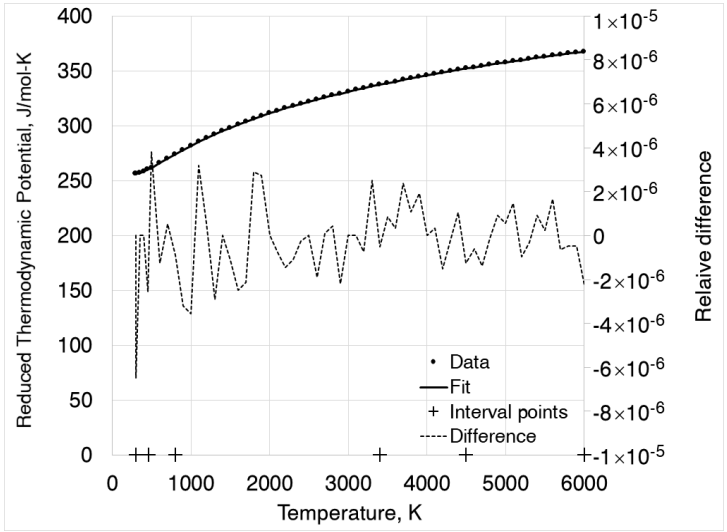
Difluorosilylene (SiF_2) – F2Si1(g)

MW = 66.0823 g/mol – $\Delta_f H^o$(298.15 K) = -587.852 kJ/mol – S^o(298.15 K) = 256.579 J/mol-K – p^o = 0.1 MPa



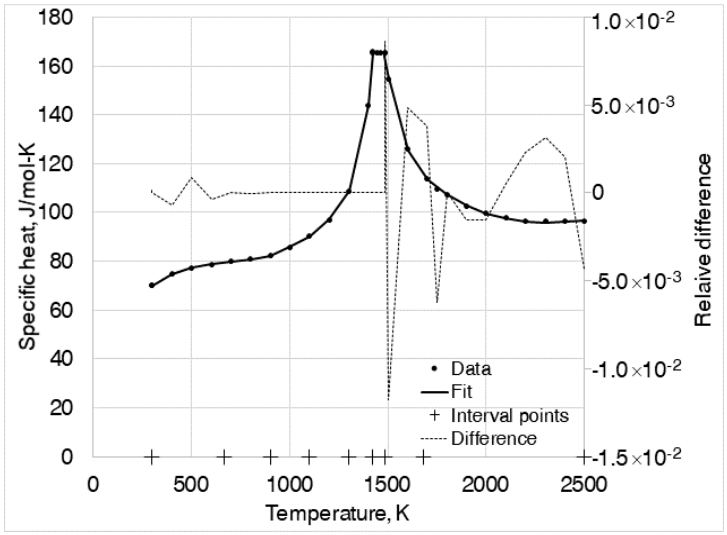


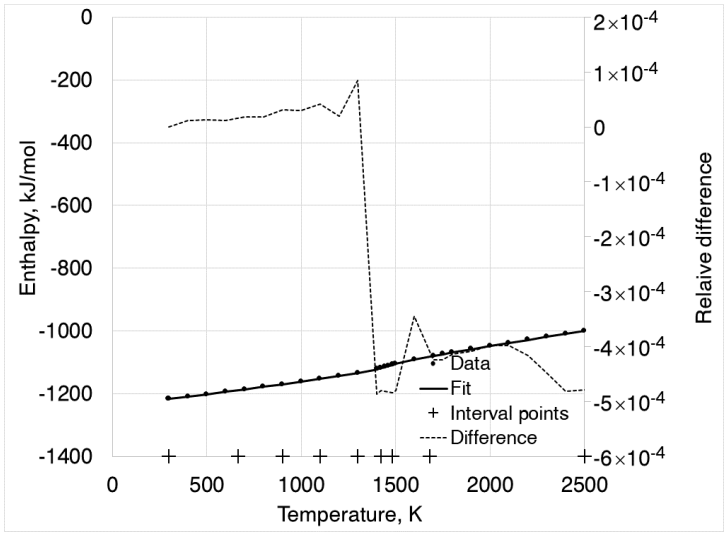


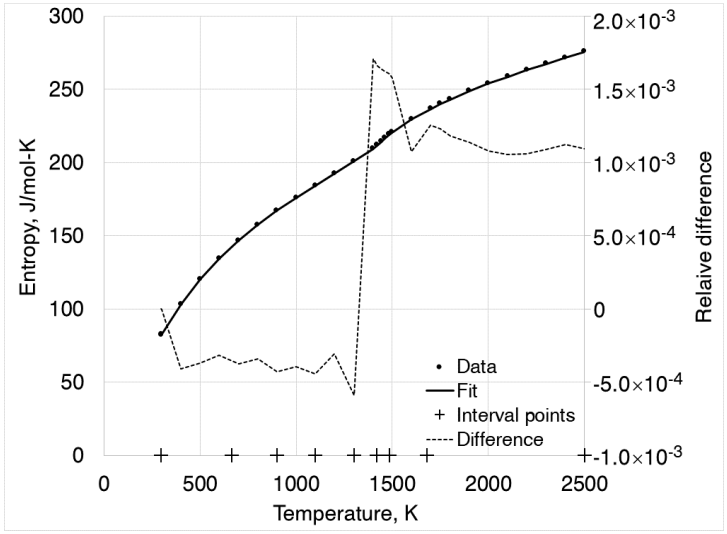


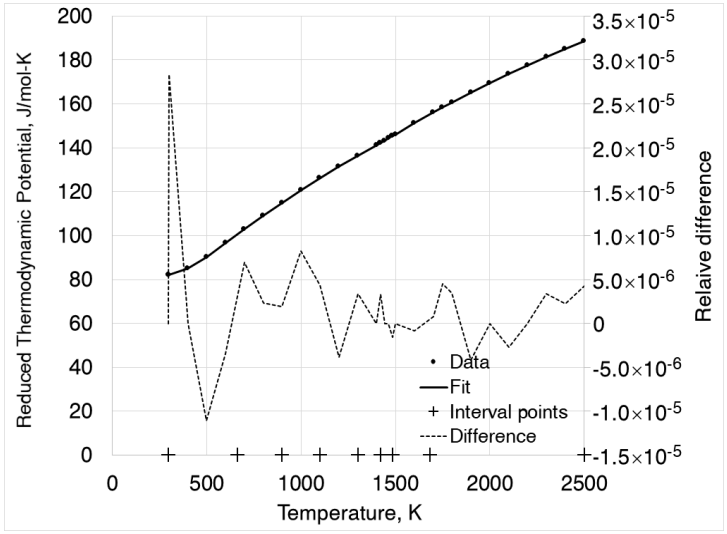
Strontium Fluoride (SrF_2) – F2Sr1(cr)

MW = 125.62 g/mol – $\Delta_f H^o$(298.15 K) = -1217.126 kJ/mol – S^o(298.15 K) = 82.123 J/mol-K – p^o = 0.1 MPa



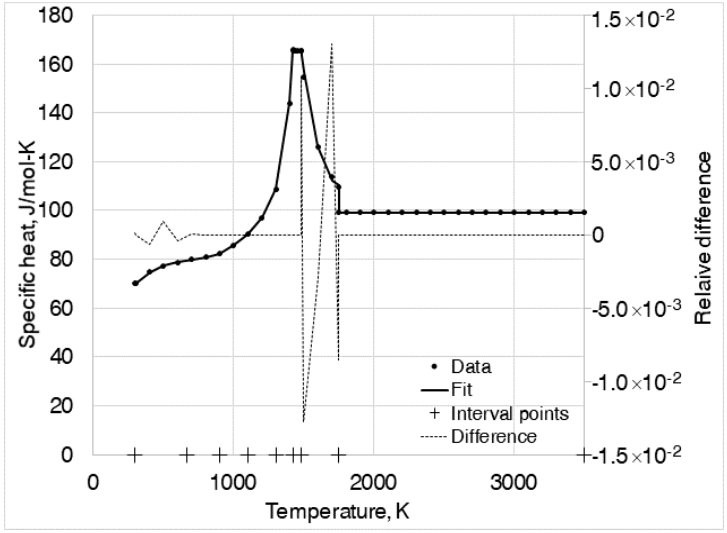


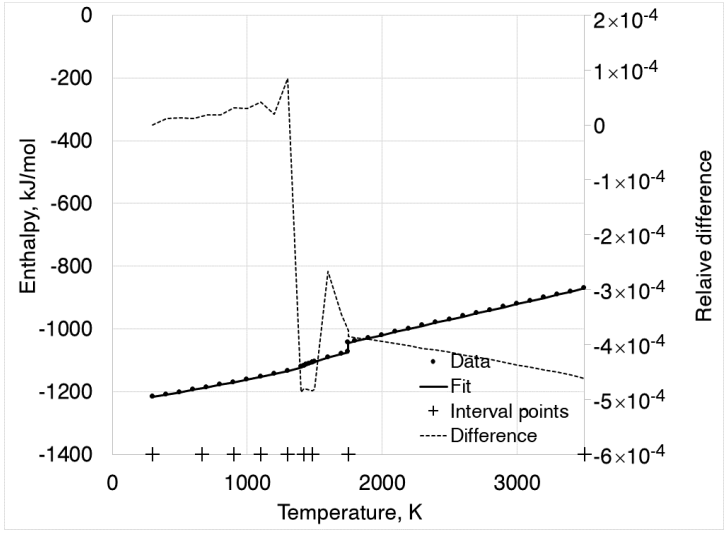


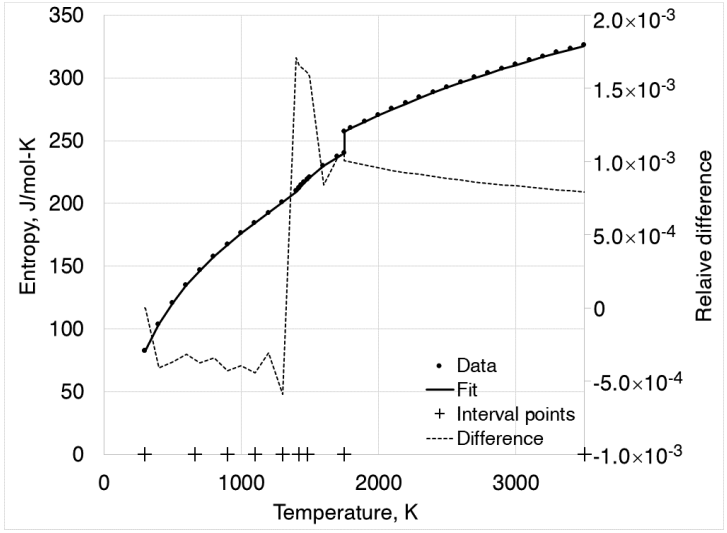


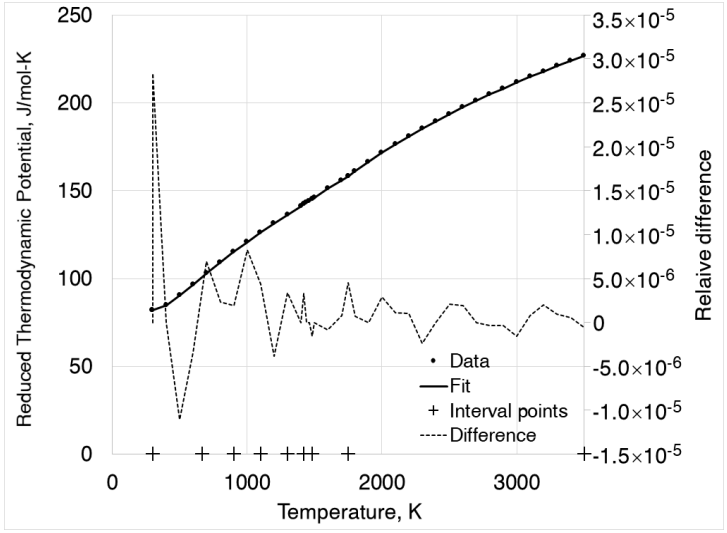
Strontium Fluoride (SrF_2) – F2Sr1(cr,l)

MW = 125.62 g/mol – $\Delta_f H^o$(298.15 K) = -1217.126 kJ/mol – S^o(298.15 K) = 82.123 J/mol-K – p^o = 0.1 MPa



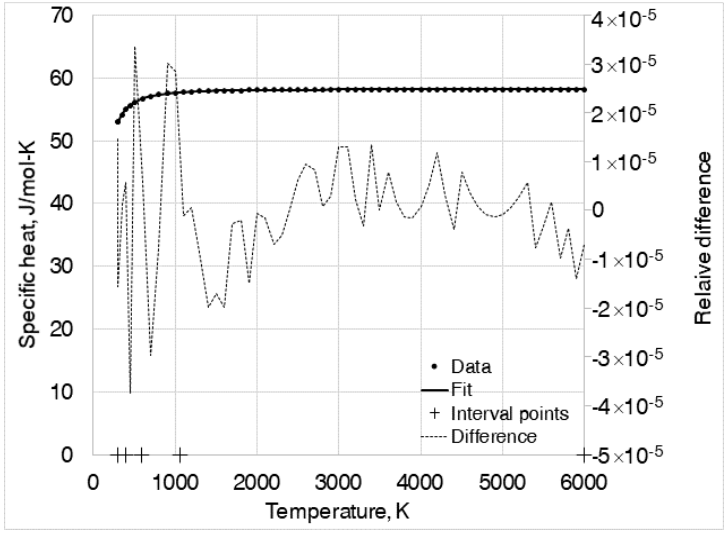


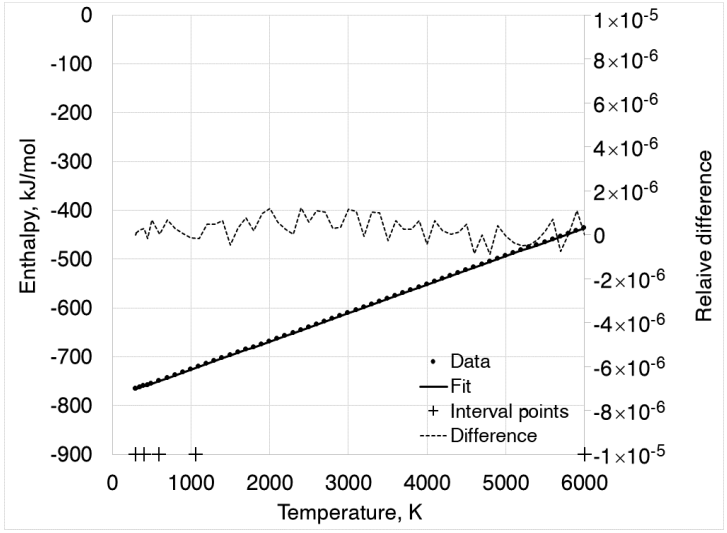


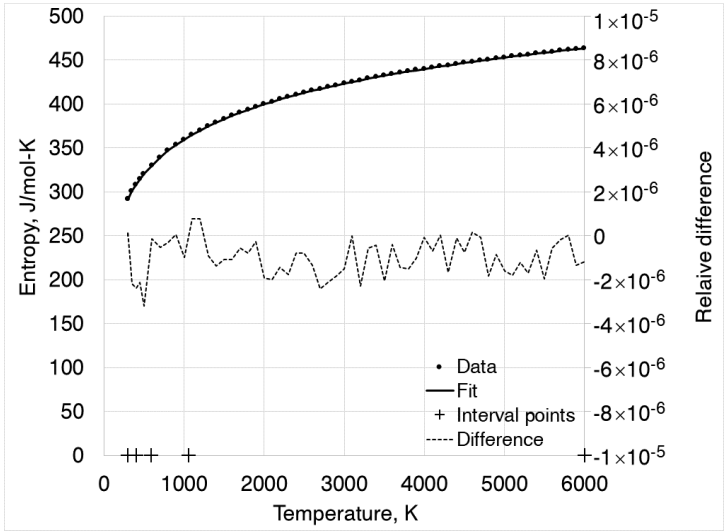


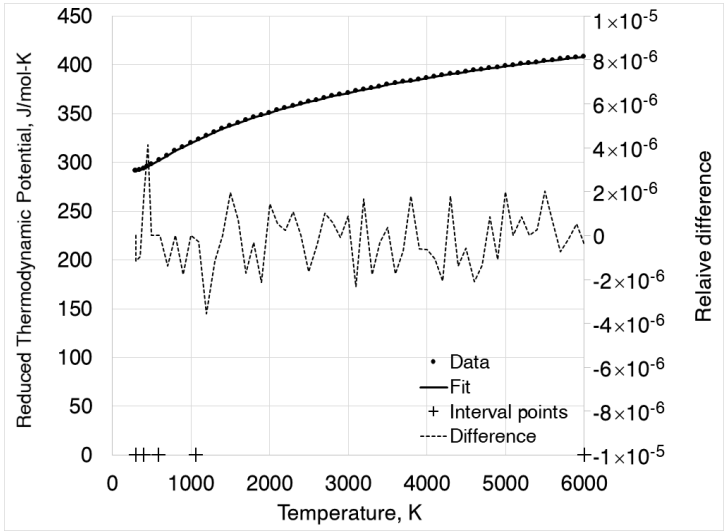
Strontium Fluoride (SrF_2) – F2Sr1(g)

MW = 125.62 g/mol – $\Delta_f H^o$(298.15 K) = -766.09 kJ/mol – S^o(298.15 K) = 291.715 J/mol-K – p^o = 0.1 MPa



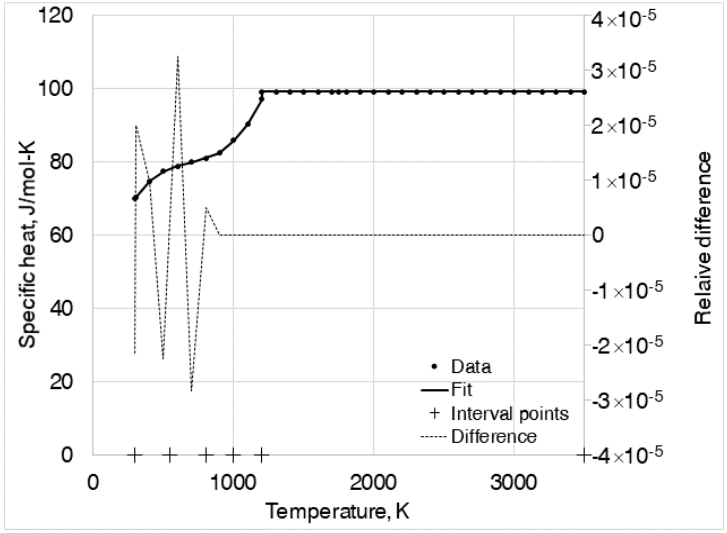


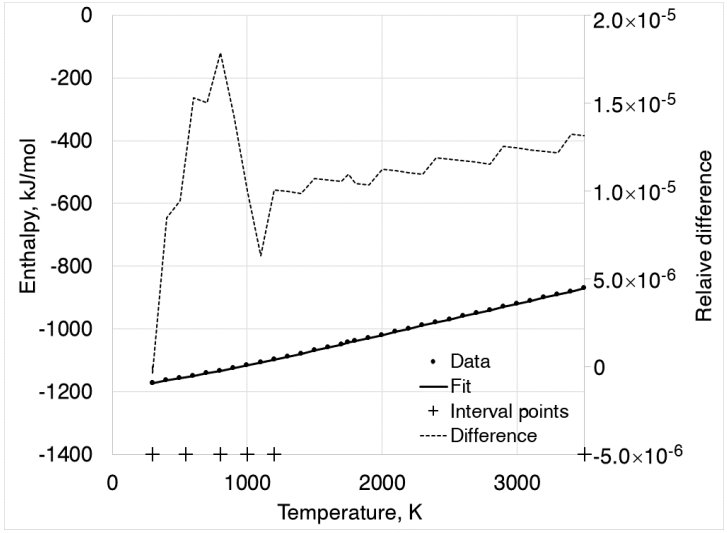


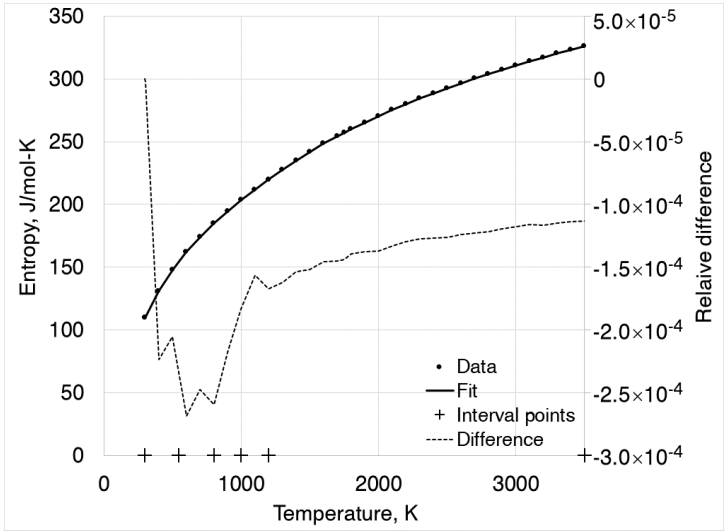


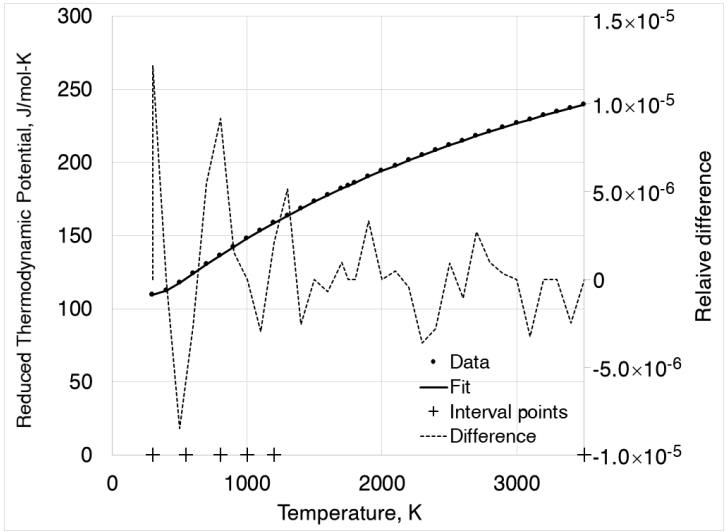
Strontium Fluoride (SrF_2) – F2Sr1(l)

MW = 125.62 g/mol – $\Delta_f H^o$(298.15 K) = -1172.043 kJ/mol – S^o(298.15 K) = 109.445 J/mol-K – p^o = 0.1 MPa



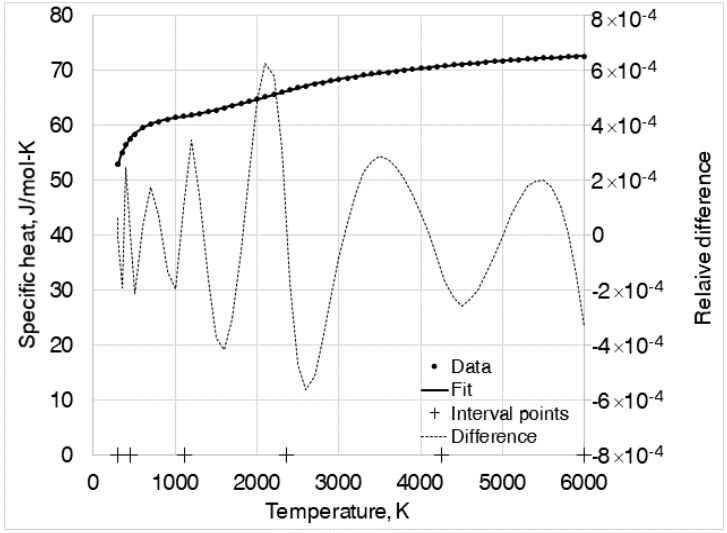


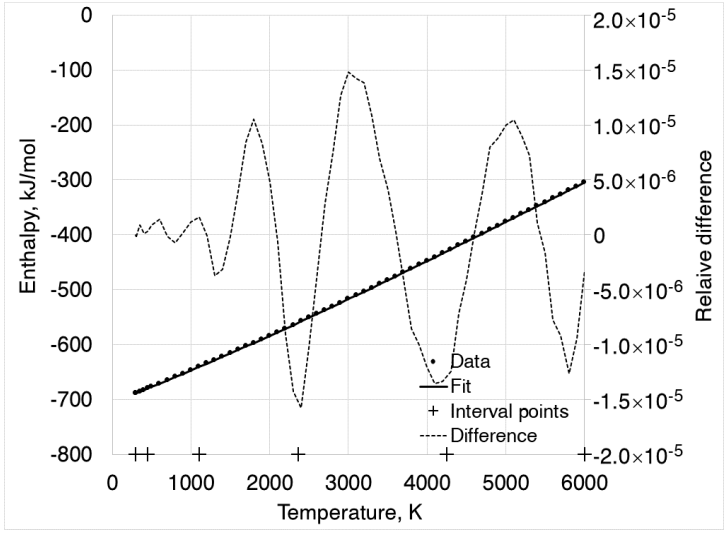


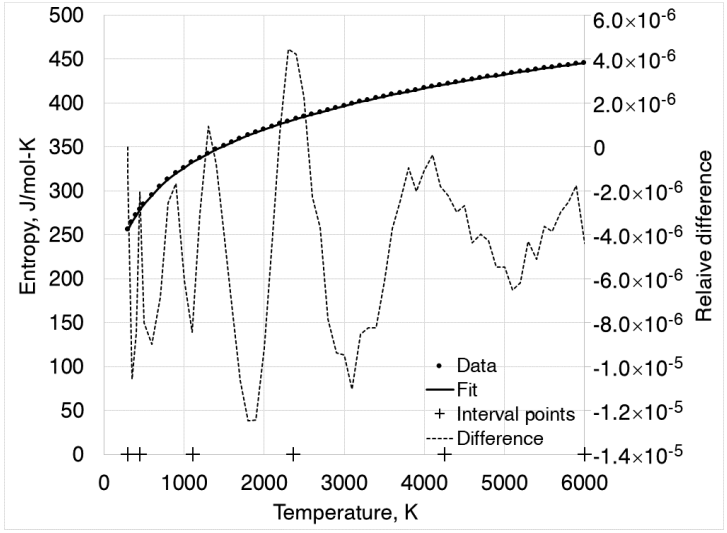


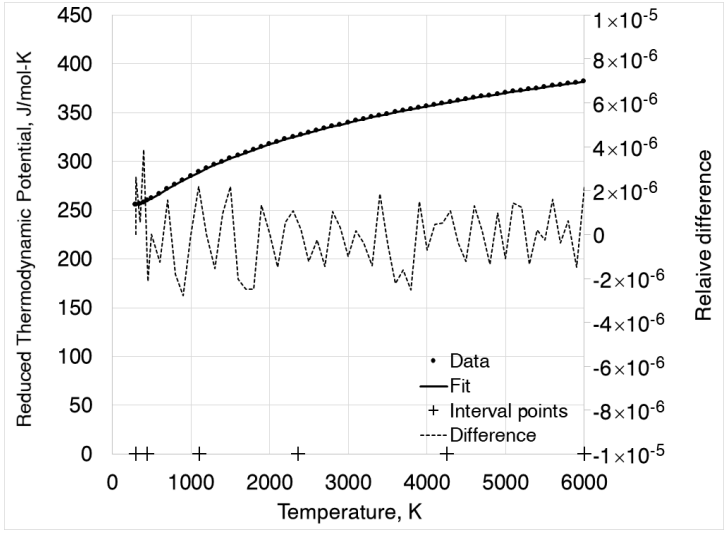
Titanium Fluoride (TiF_2) – F2Ti1(g)

MW = 85.88 g/mol – $\Delta_f H^o$(298.15 K) = -688.268 kJ/mol – S^o(298.15 K) = 255.69 J/mol-K – p^o = 0.1 MPa



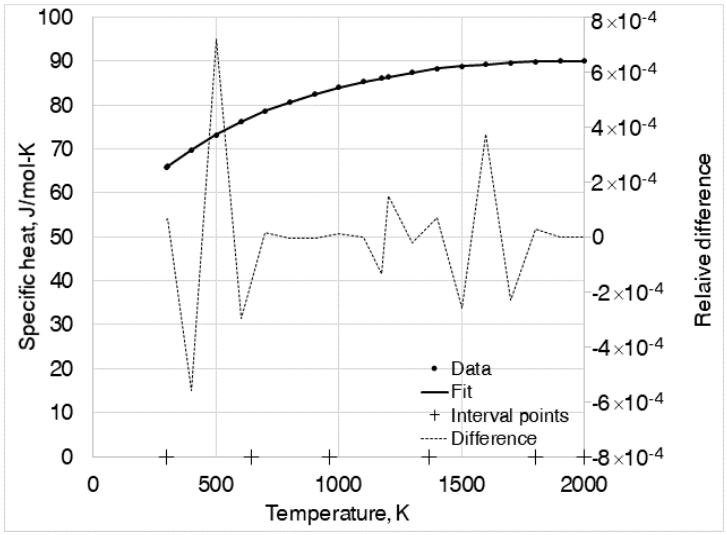


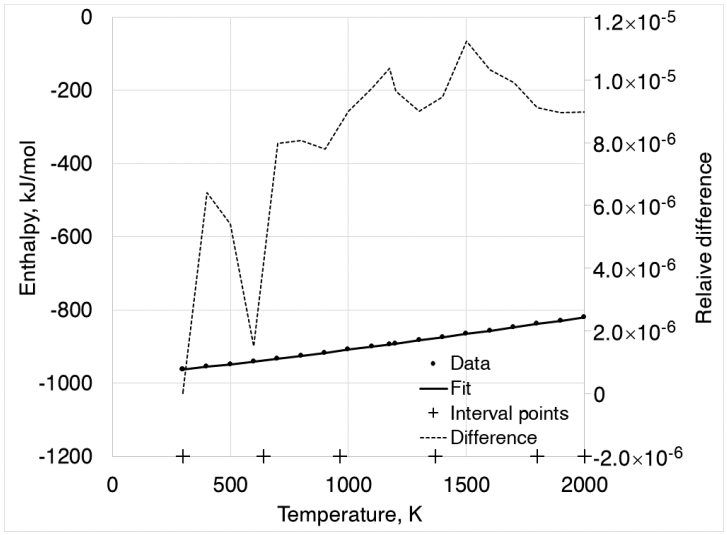


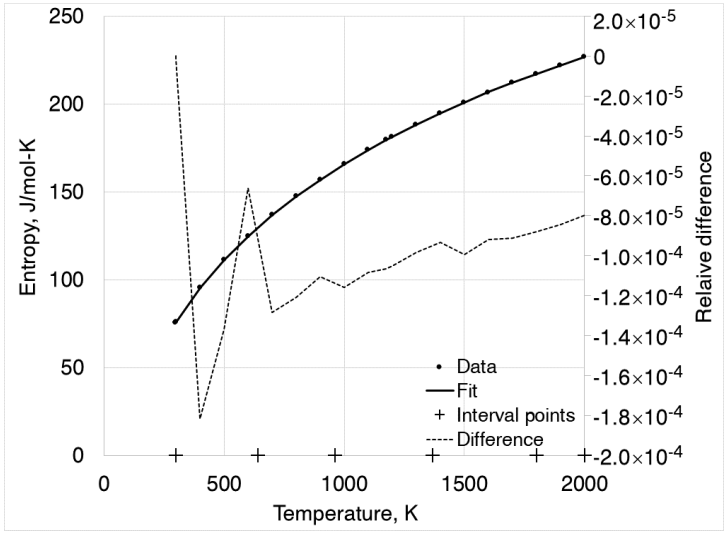


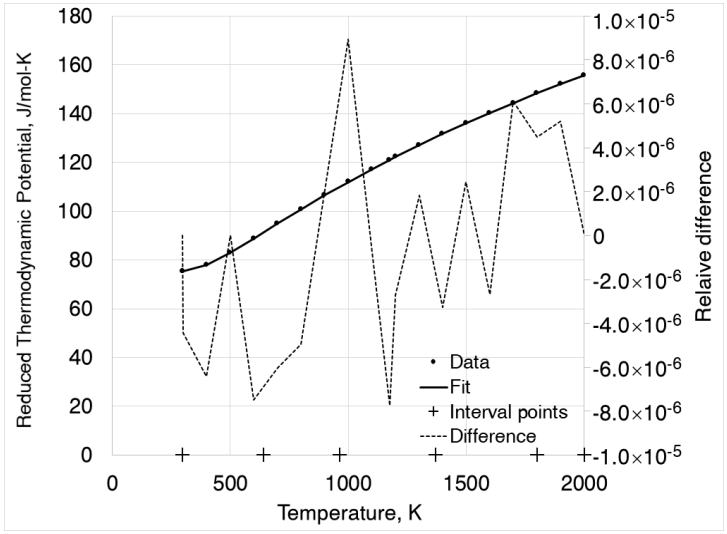
Zirconium Fluoride (ZrF_2) – F2Zr1(cr)

MW = 129.22 g/mol – $\Delta_f H^o$(298.15 K) = -962.32 kJ/mol – S^o(298.15 K) = 75.312 J/mol-K – p^o = 0.1 MPa



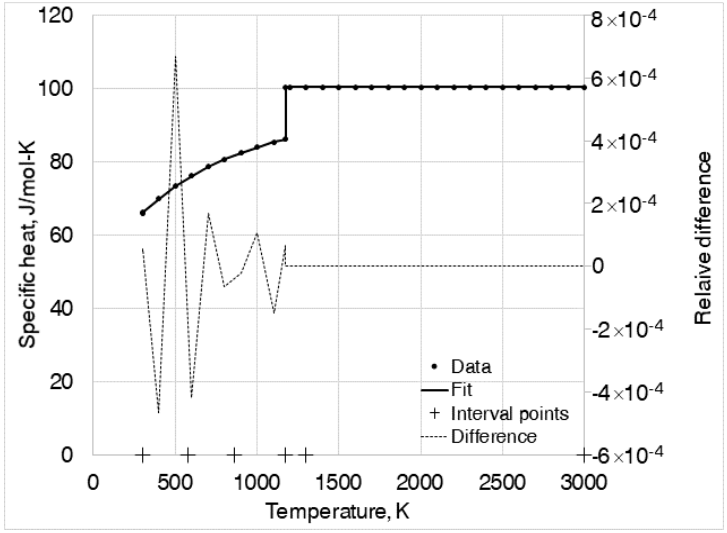


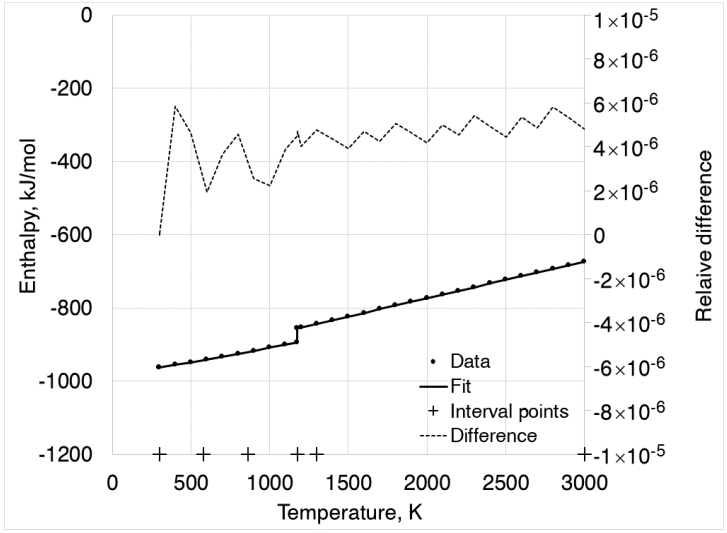


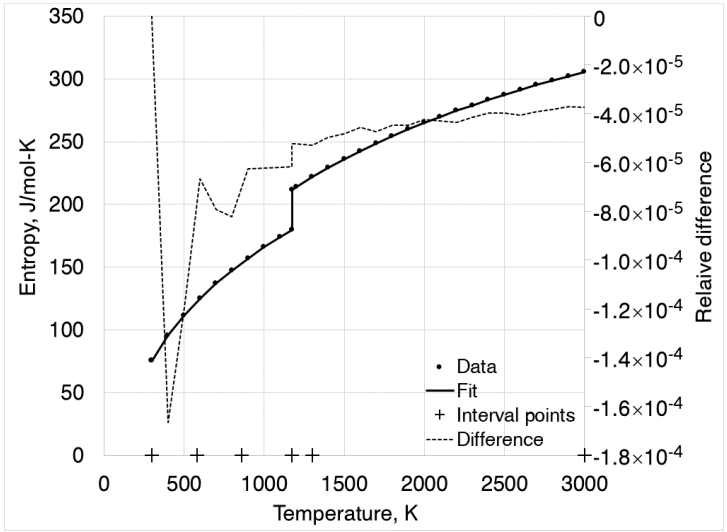


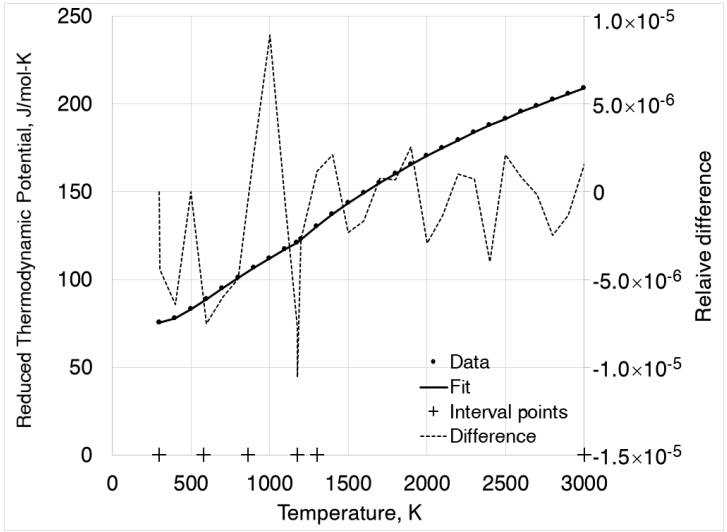
Zirconium Fluoride (ZrF_2) – F2Zr1(cr,l)

MW = 129.22 g/mol – $\Delta_f H^o$(298.15 K) = -962.32 kJ/mol – S^o(298.15 K) = 75.312 J/mol-K – p^o = 0.1 MPa



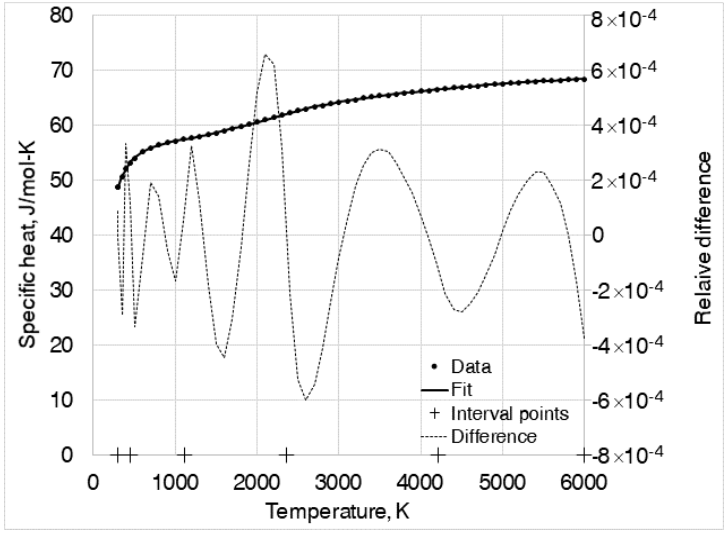


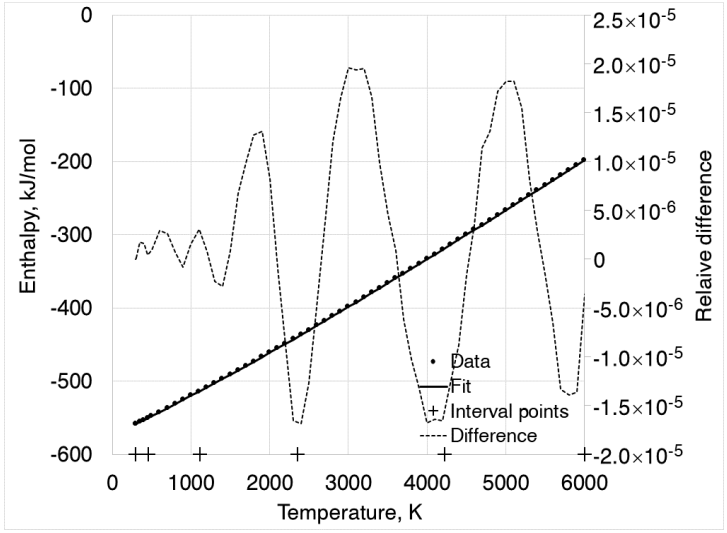


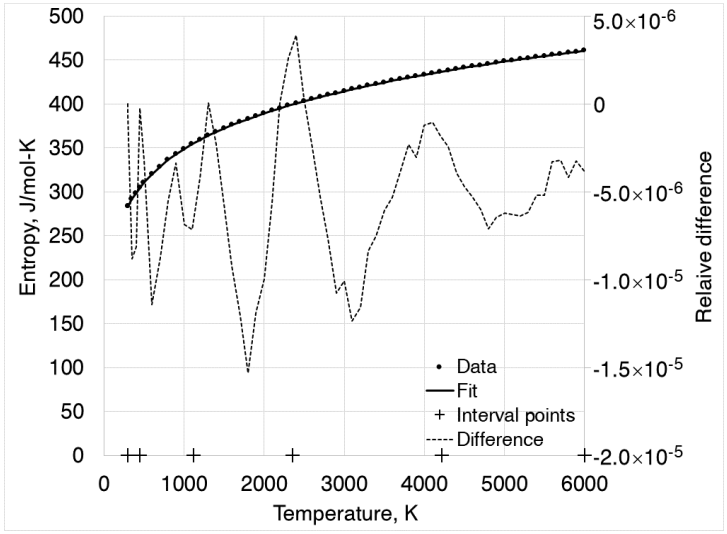


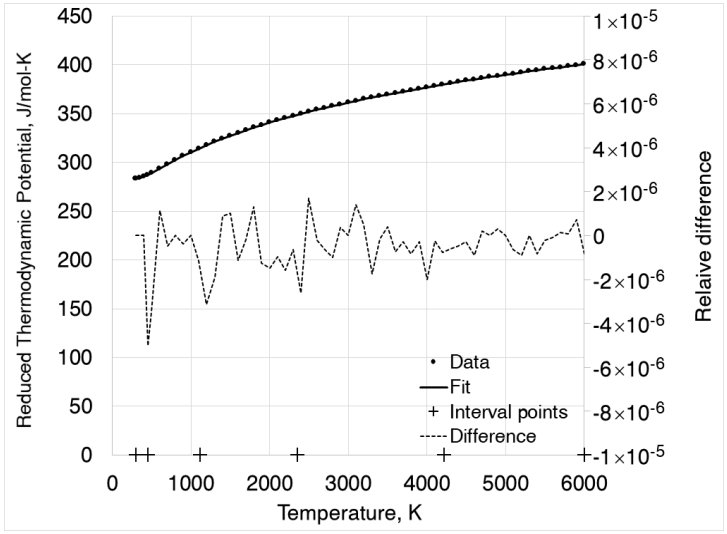
Zirconium Fluoride (ZrF_2) – F2Zr1(g)

MW = 129.22 g/mol – $\Delta_f H^o$(298.15 K) = -558.146 kJ/mol – S^o(298.15 K) = 283.434 J/mol-K – p^o = 0.1 MPa



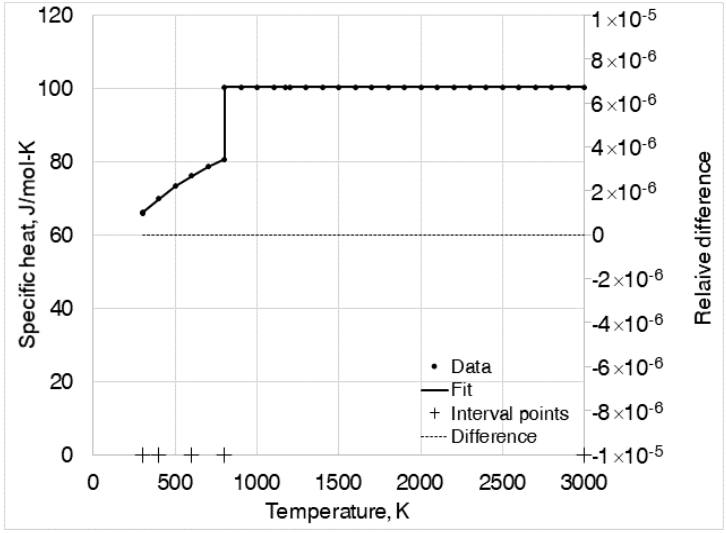


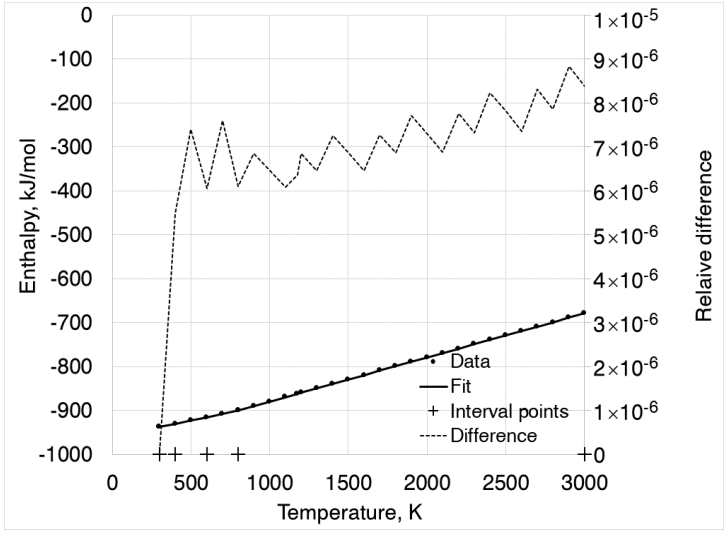


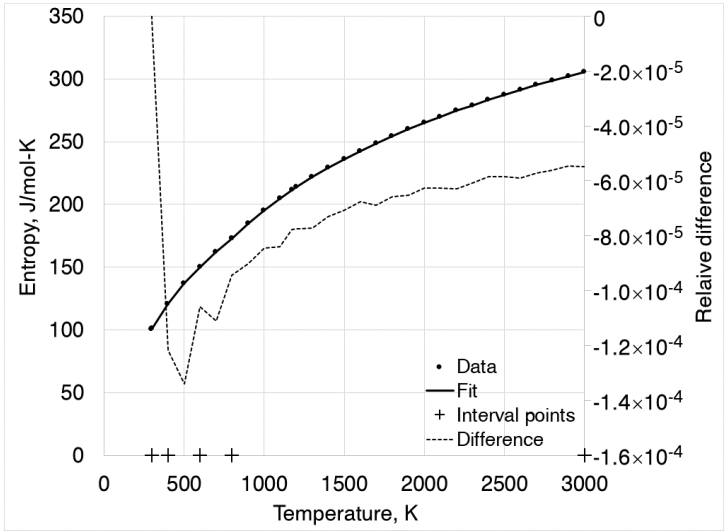


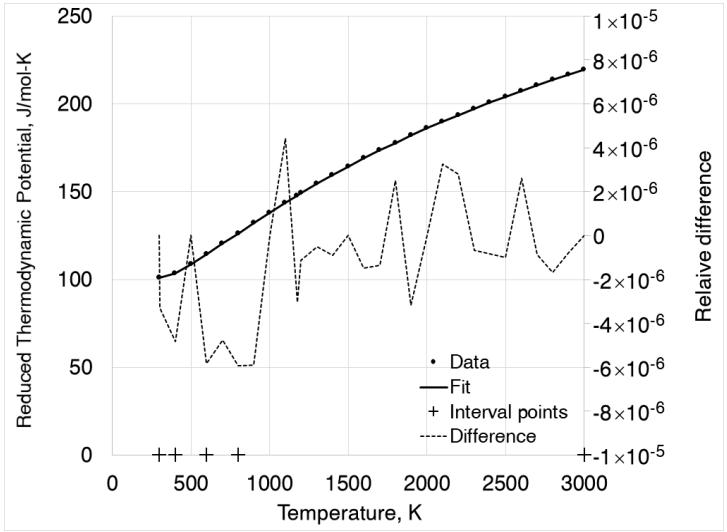
Zirconium Fluoride (ZrF_2) – F2Zr1(l)

MW = 129.22 g/mol – $\Delta_f H^o$(298.15 K) = -936.951 kJ/mol – S^o(298.15 K) = 100.846 J/mol-K – p^o = 0.1 MPa



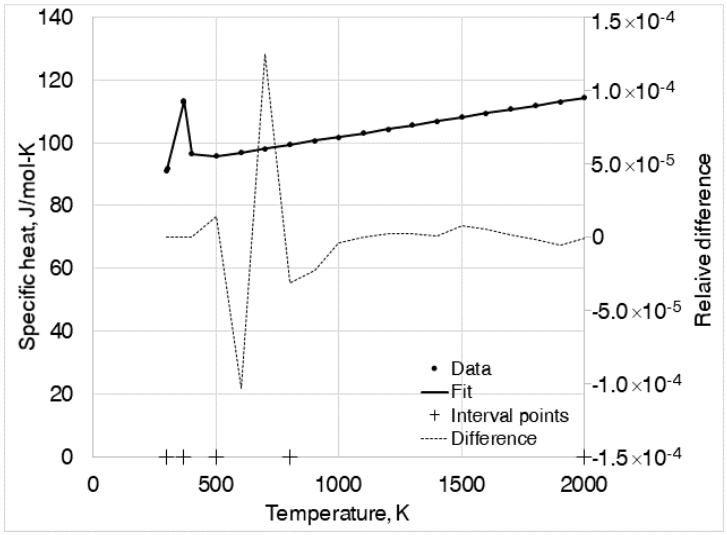


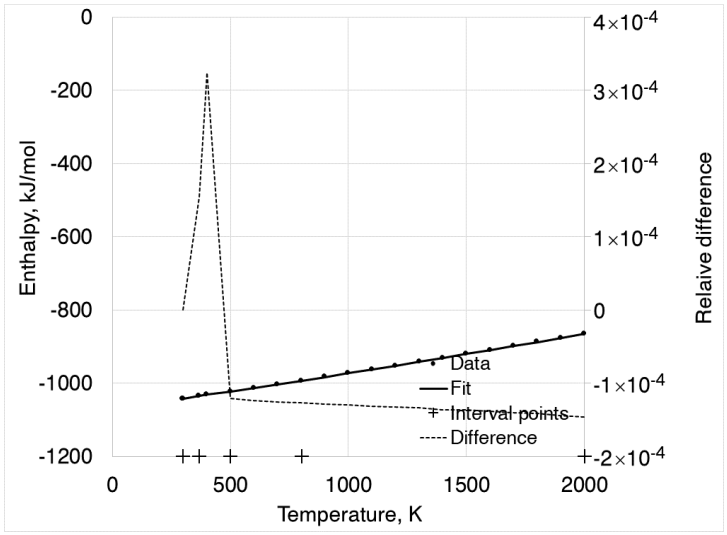


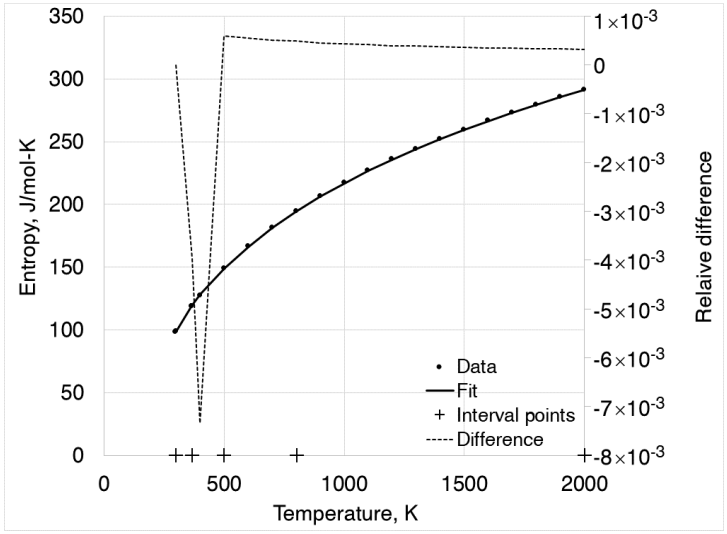


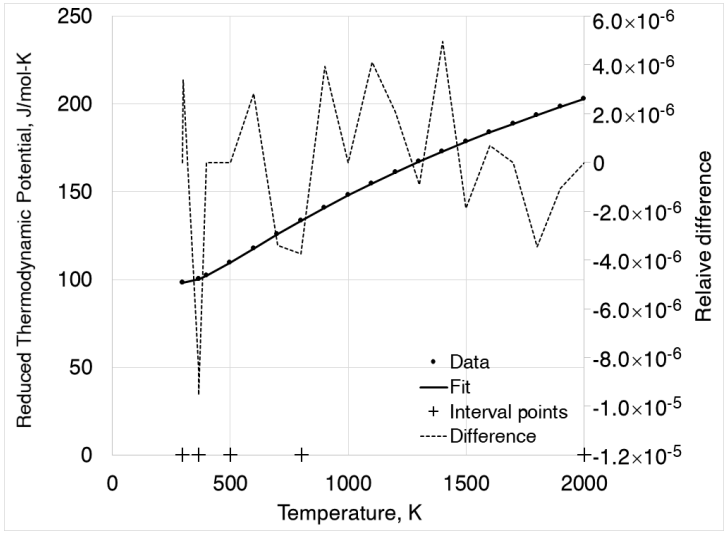
Iron Fluoride (FeF_3) – F3Fe1(cr)

MW = 112.842 g/mol – $\Delta_f H^o$(298.15 K) = -1041.816 kJ/mol – S^o(298.15 K) = 98.324 J/mol-K – p^o = 0.1 MPa



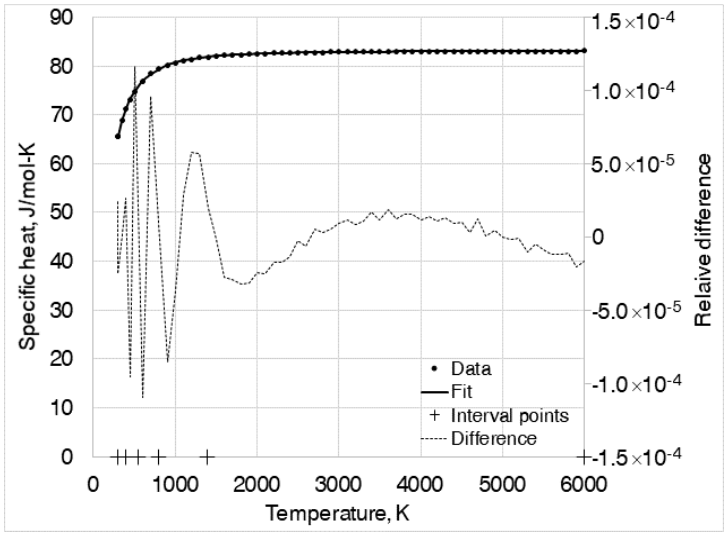


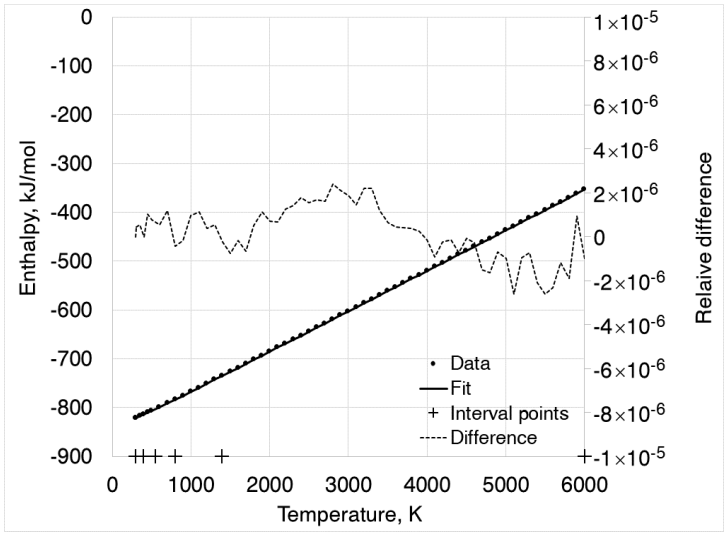


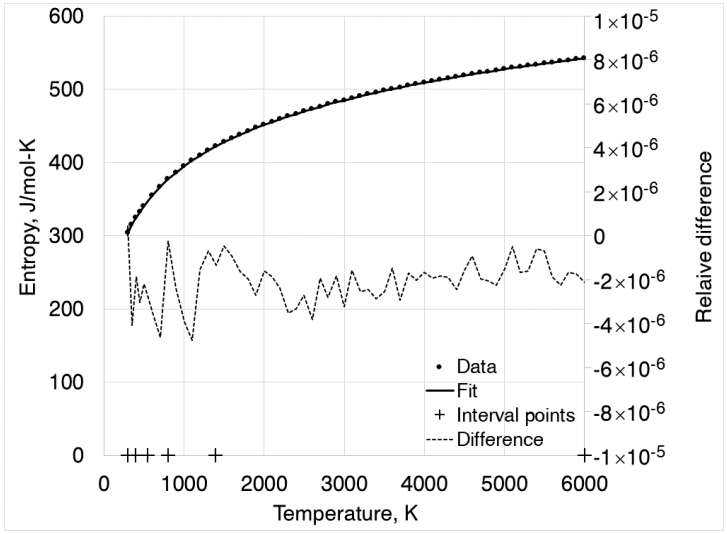


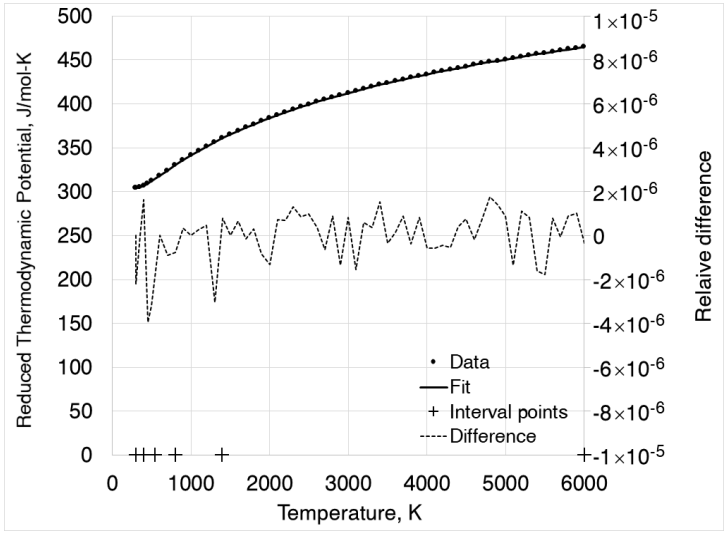
Iron Fluoride (FeF_3) – F3Fe1(g)

MW = 112.842 g/mol – $\Delta_f H^o$(298.15 K) = -820.901 kJ/mol – S^o(298.15 K) = 304.236 J/mol-K – p^o = 0.1 MPa



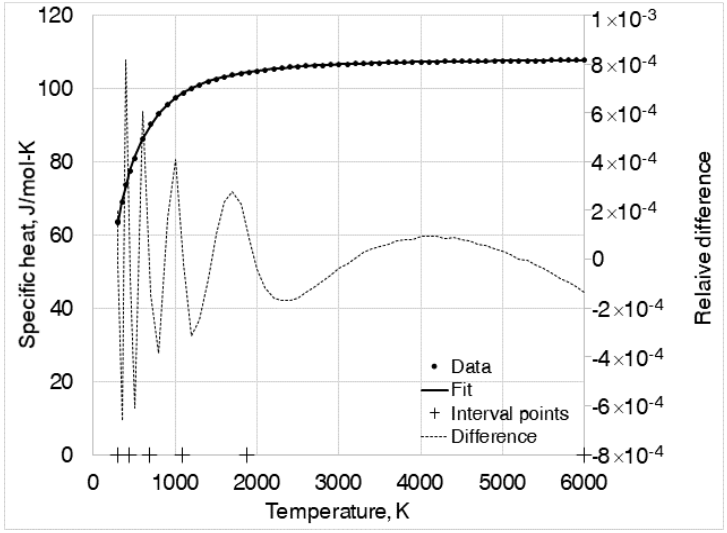


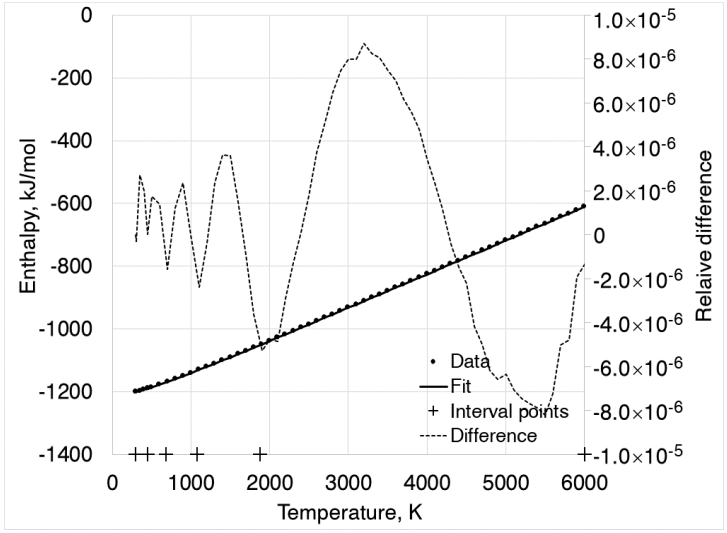


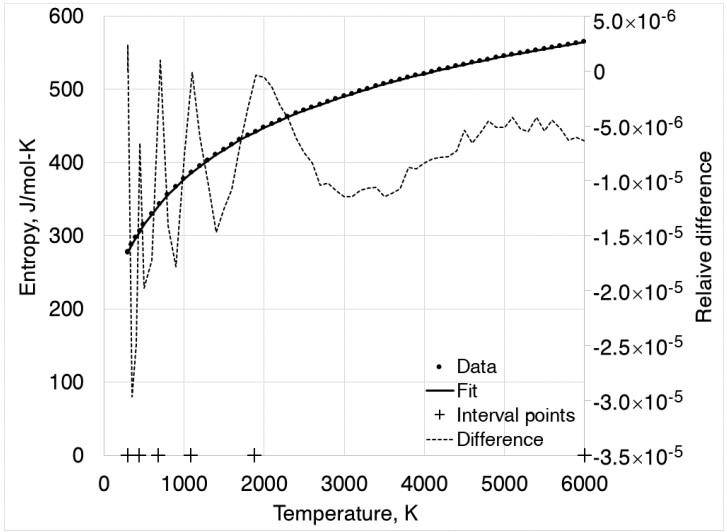


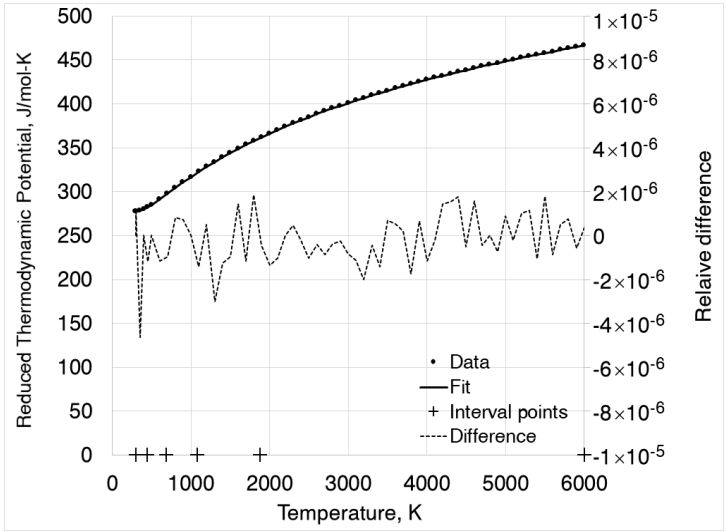
Trifluorosilane ($SiHF_3$) – F3H1Si1(g)

MW = 86.0886 g/mol – $\Delta_f H^o$(298.15 K) = -1200.808 kJ/mol – S^o(298.15 K) = 277.268 J/mol-K – p^o = 0.1 MPa



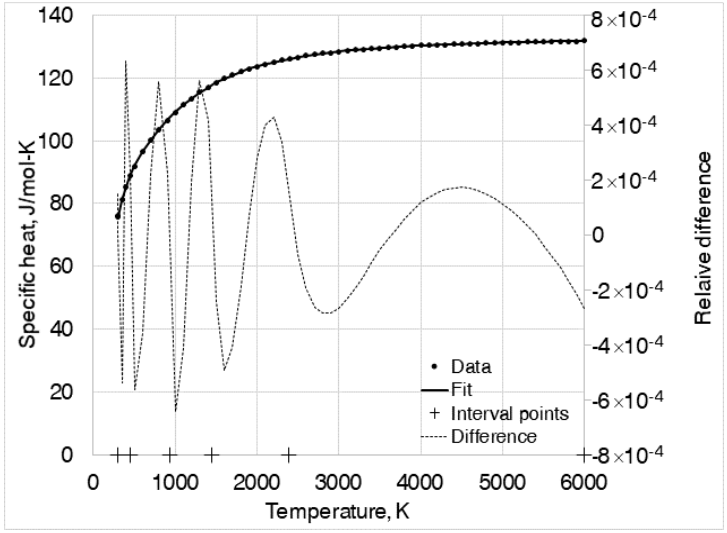


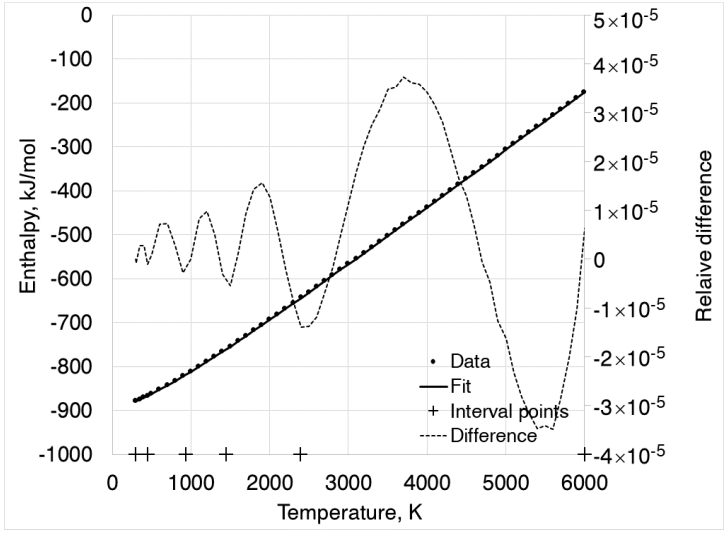


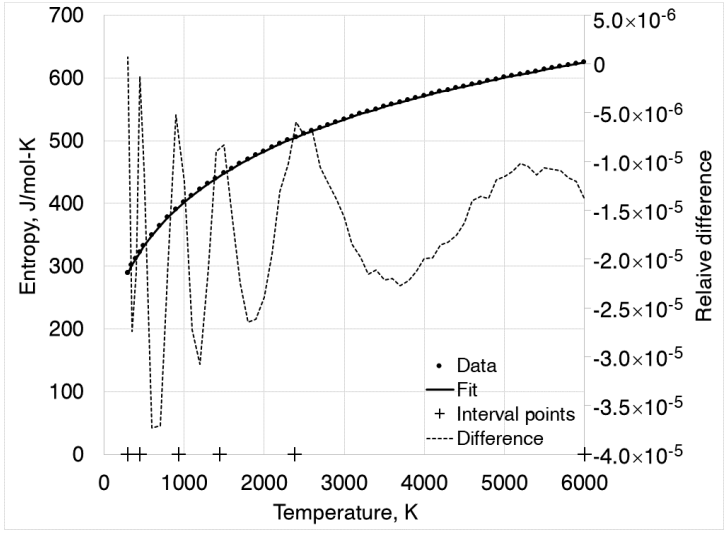


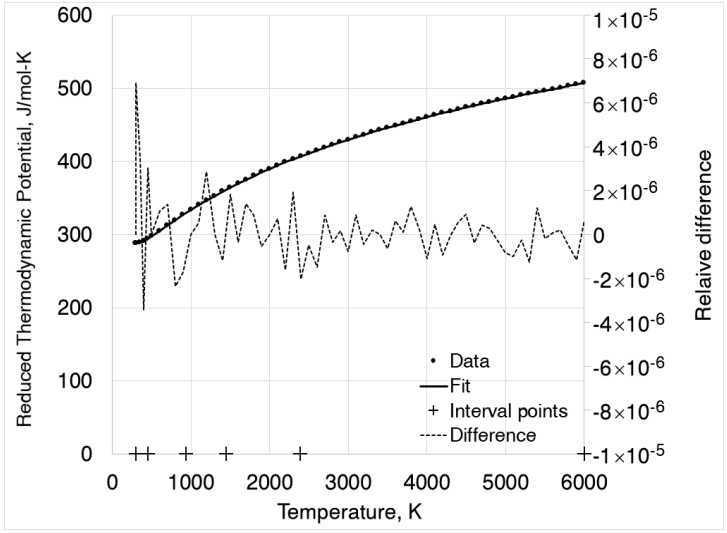
Hydrogen Fluoride (H_3F_3) – F3H3(g)

MW = 60.019 g/mol – $\Delta_f H^o$(298.15 K) = -879.1 kJ/mol – S^o(298.15 K) = 288.458 J/mol-K – p^o = 0.1 MPa



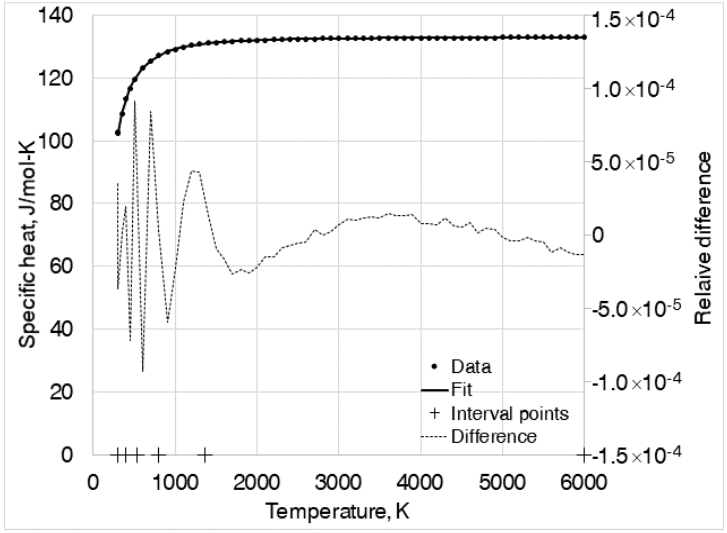


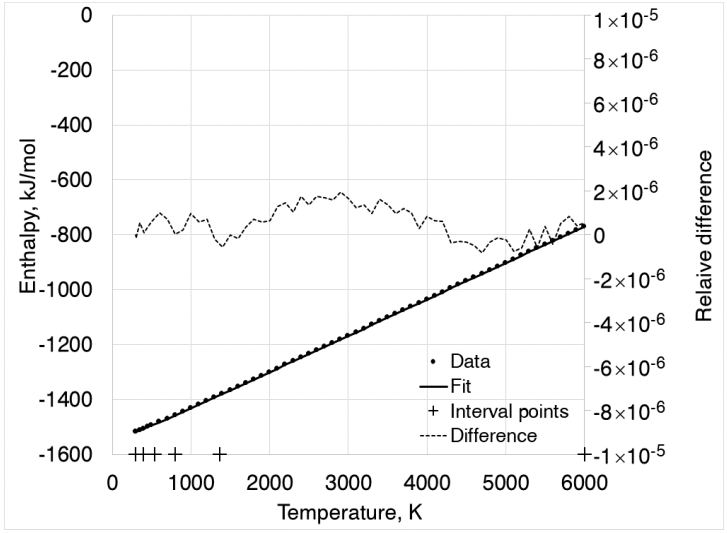


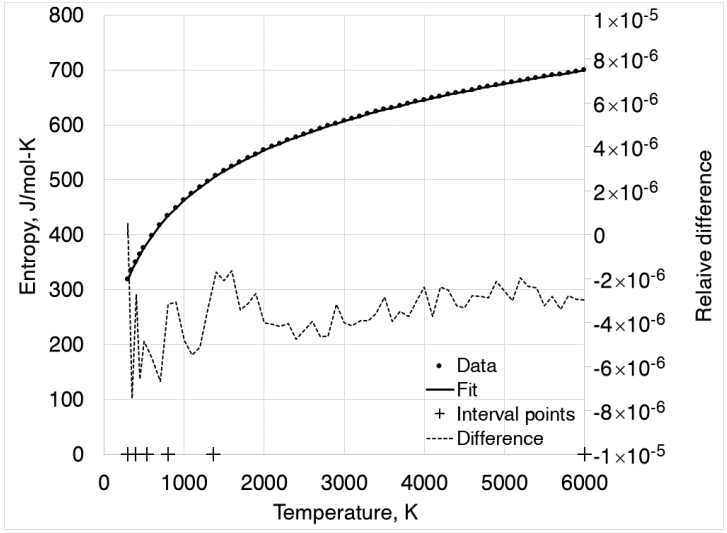


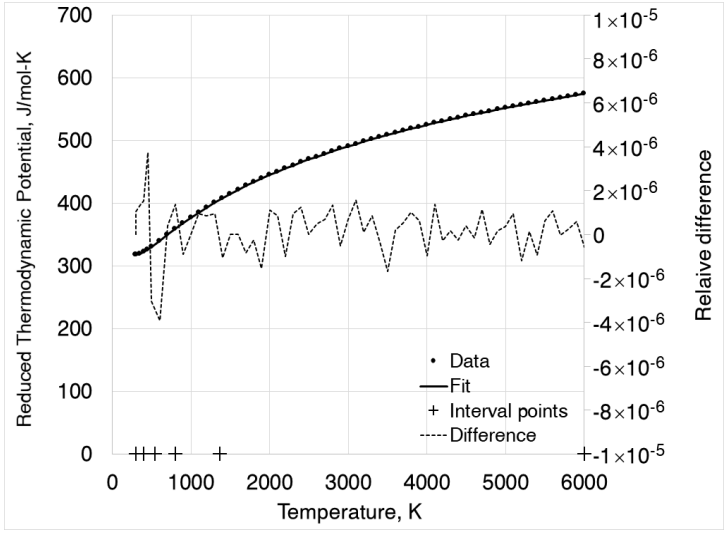
Lithium Fluoride ($(LiF)_3$) – F3Li3(g)

MW = 77.818 g/mol – $\Delta_f H^o$(298.15 K) = -1517.202 kJ/mol – S^o(298.15 K) = 318.087 J/mol-K – p^o = 0.1 MPa



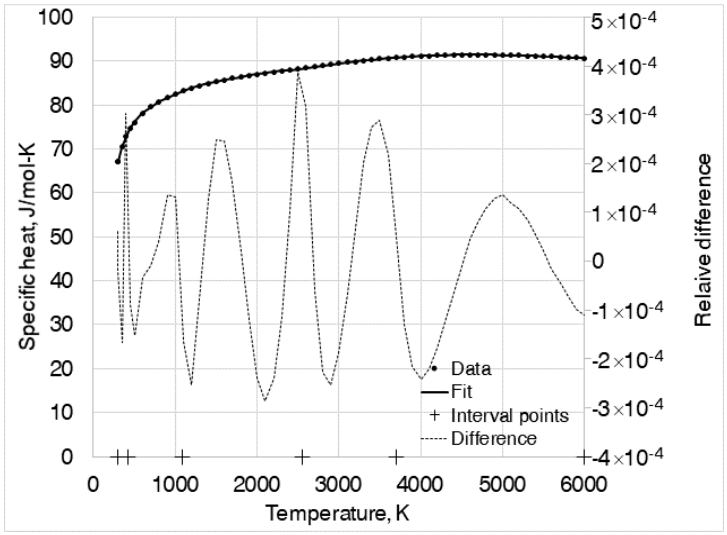


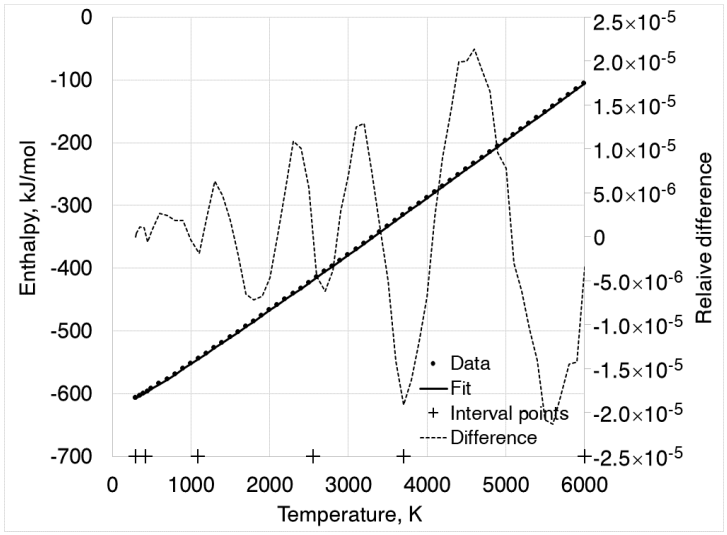


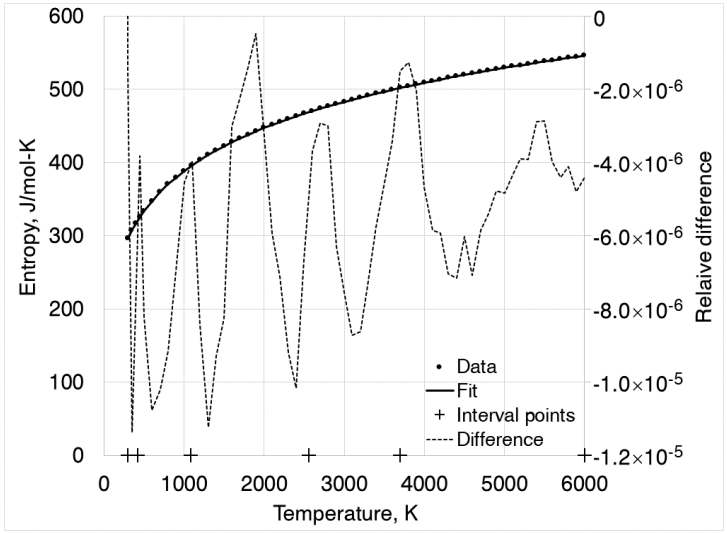


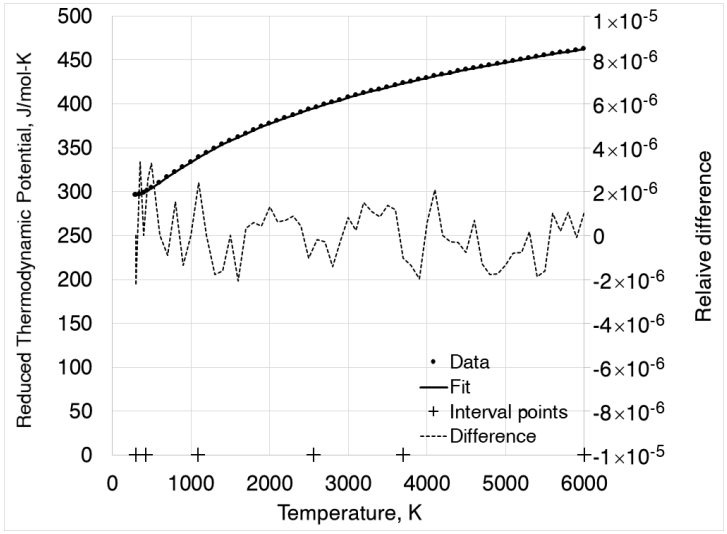
Molybdenum Fluoride (MoF_3) – F3Mo1(g)

MW = 152.94 g/mol – $\Delta_f H^o$(298.15 K) = -607.098 kJ/mol – S^o(298.15 K) = 295.968 J/mol-K – p^o = 0.1 MPa



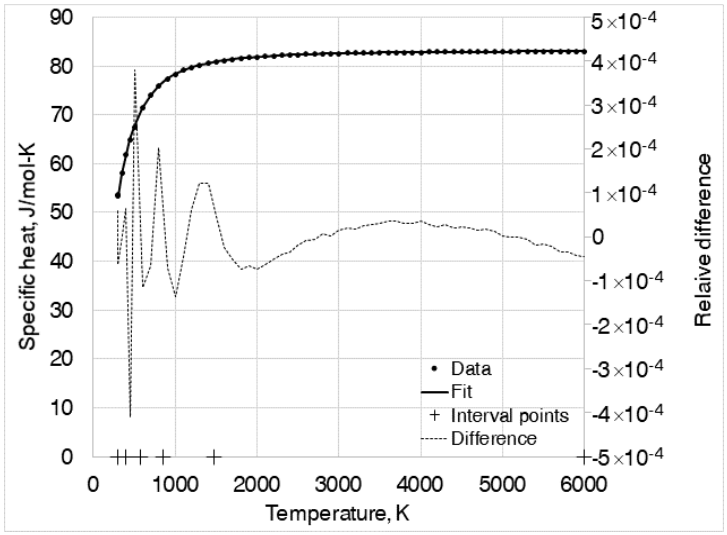


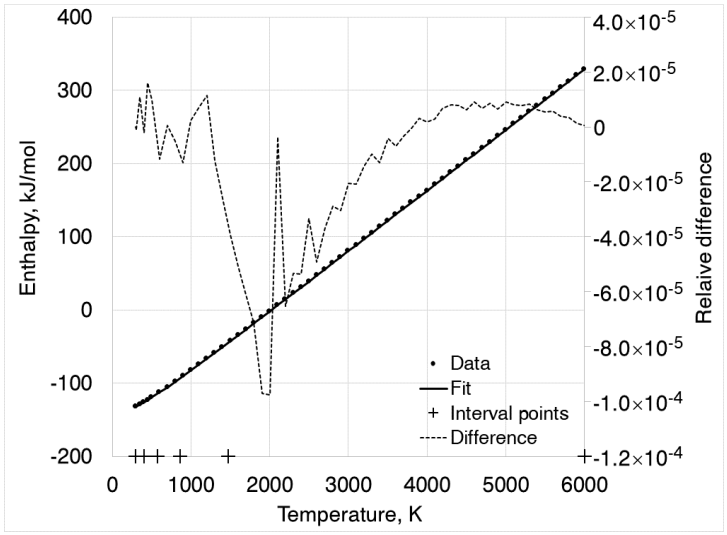


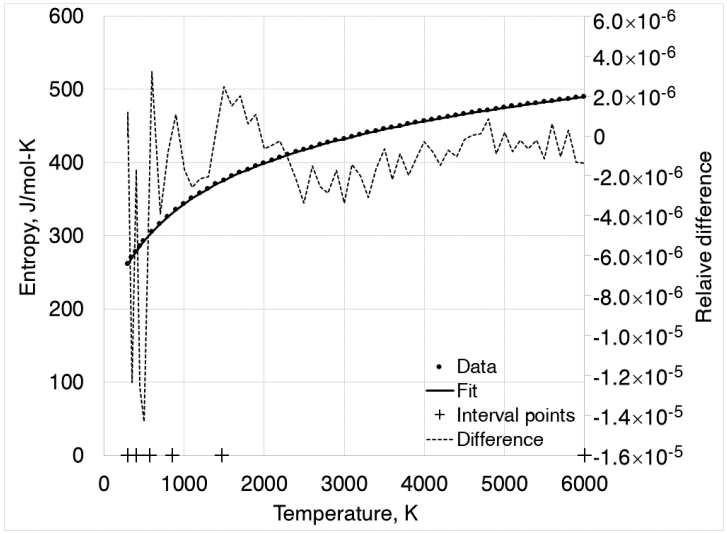


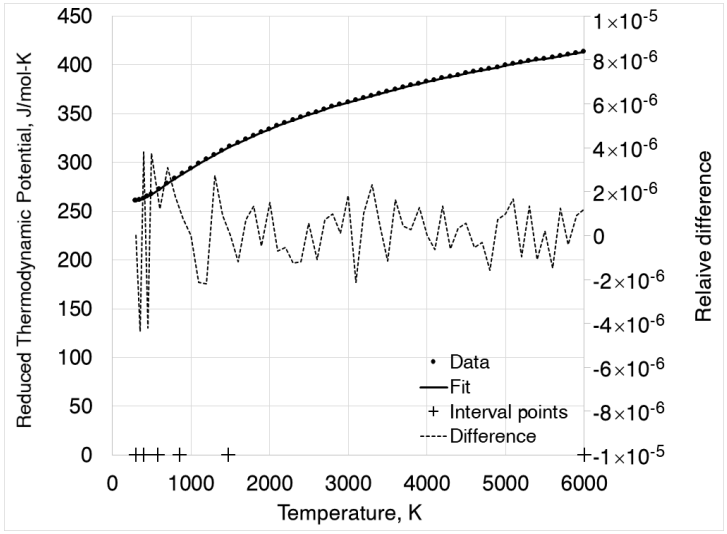
Nitrogen Fluoride (NF_3) – F3N1(g)

MW = 71.0019 g/mol – $\Delta_f H^o$(298.15 K) = -132.089 kJ/mol – S^o(298.15 K) = 260.773 J/mol-K – p^o = 0.1 MPa



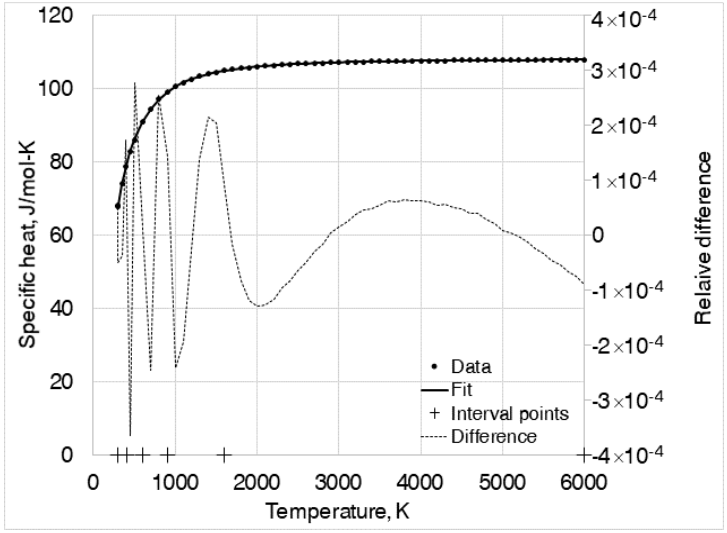


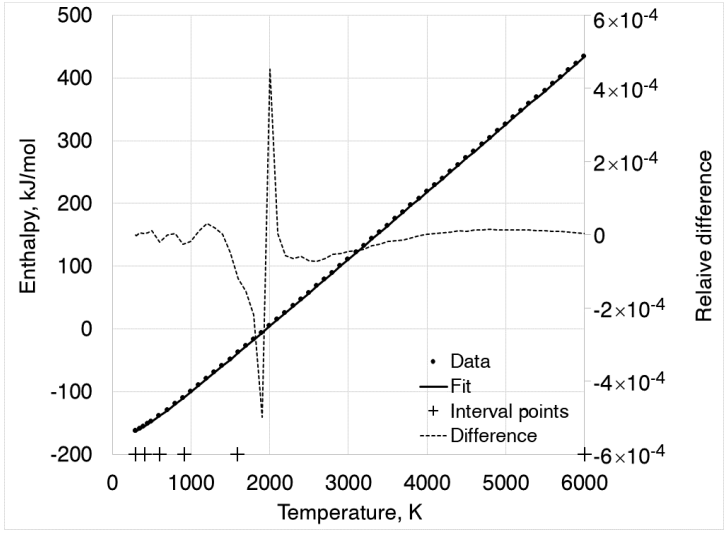


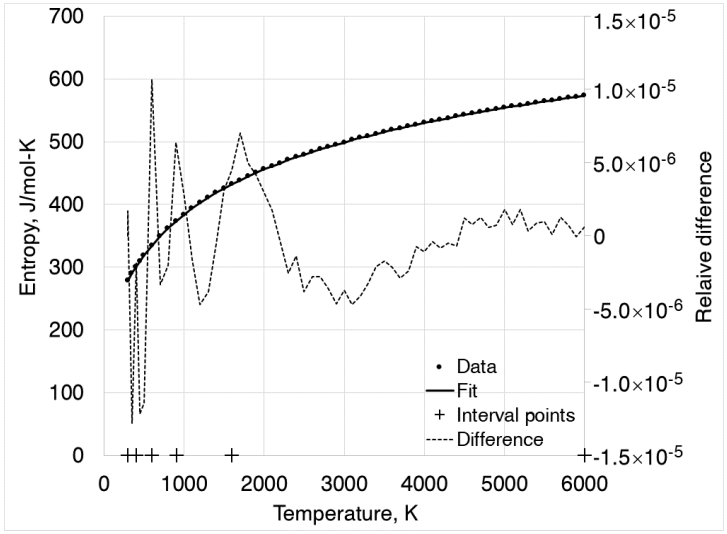


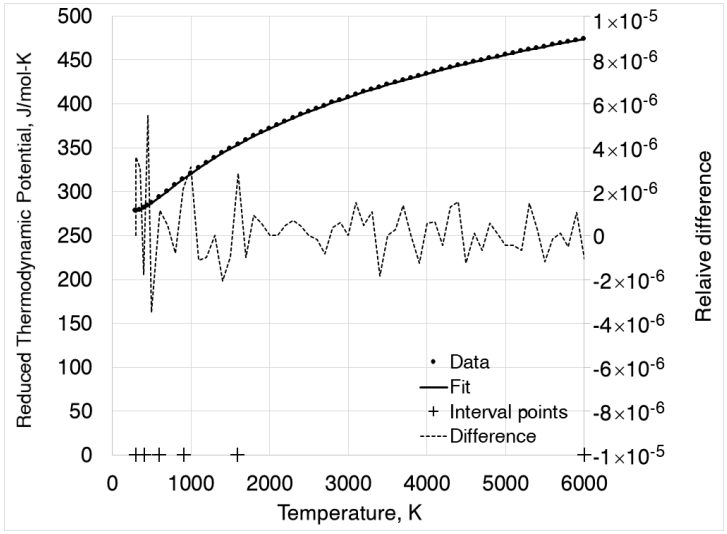
Nitrogen Fluoride Oxide (NF_3O) – F3N1O1(g)

MW = 87.0013 g/mol – $\Delta_f H^o$(298.15 K) = -163.176 kJ/mol – S^o(298.15 K) = 278.49 J/mol-K – p^o = 0.1 MPa



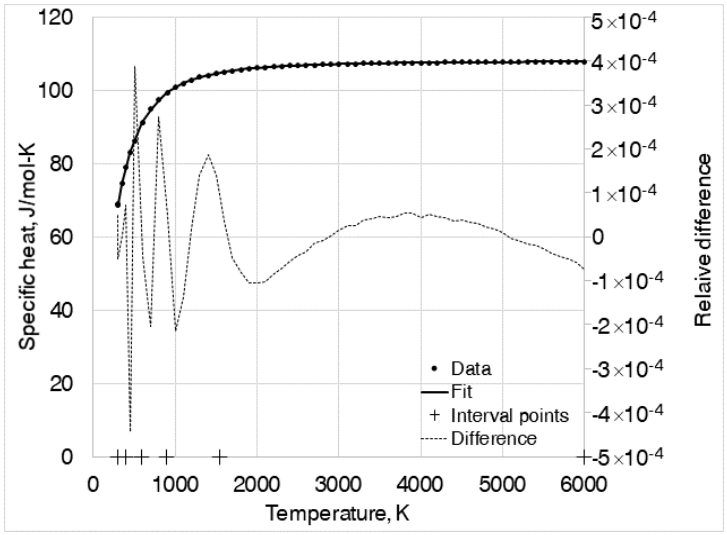


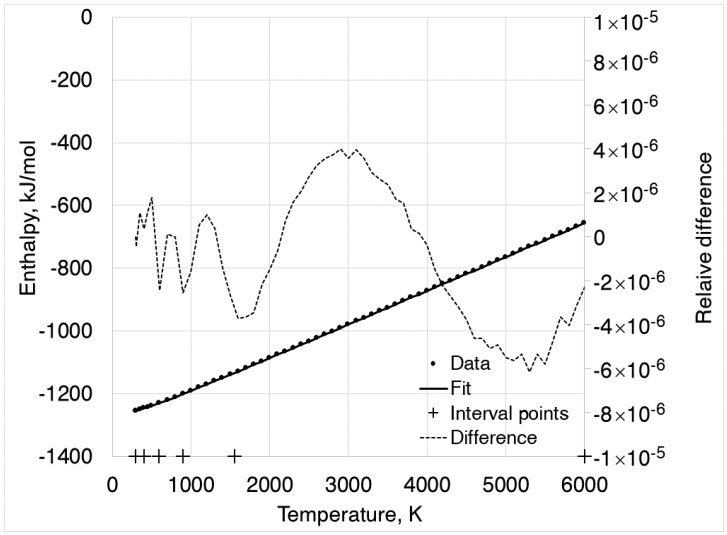


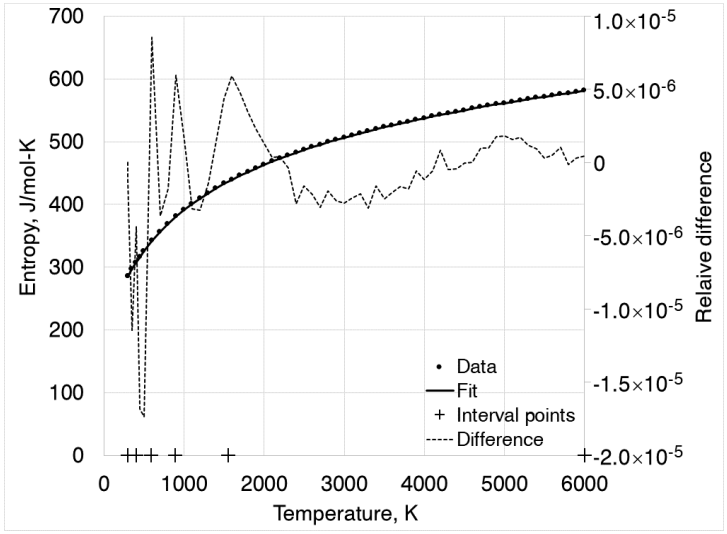


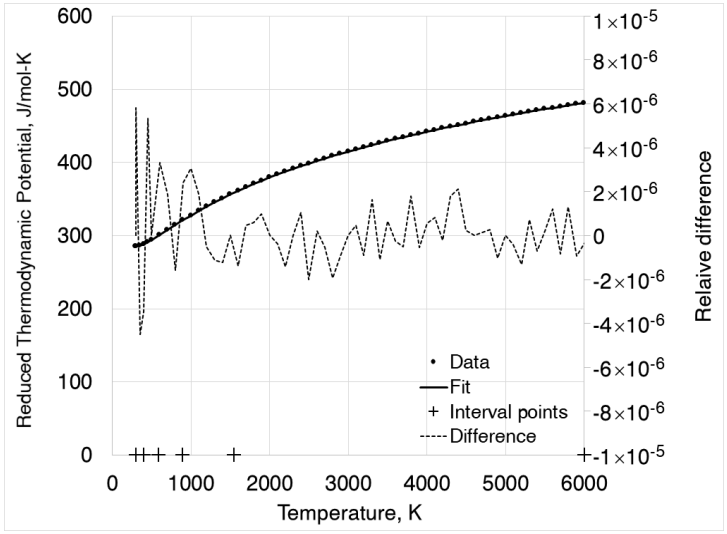
Phosphoryl Fluoride (POF_3) – F3O1P1(g)

MW = 103.9684 g/mol – $\Delta_f H^o$(298.15 K) = -1254.25 kJ/mol – S^o(298.15 K) = 285.413 J/mol-K – p^o = 0.1 MPa



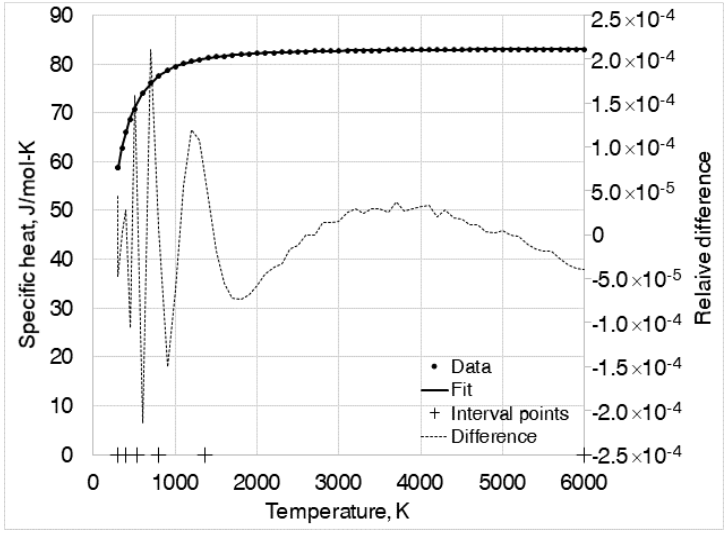


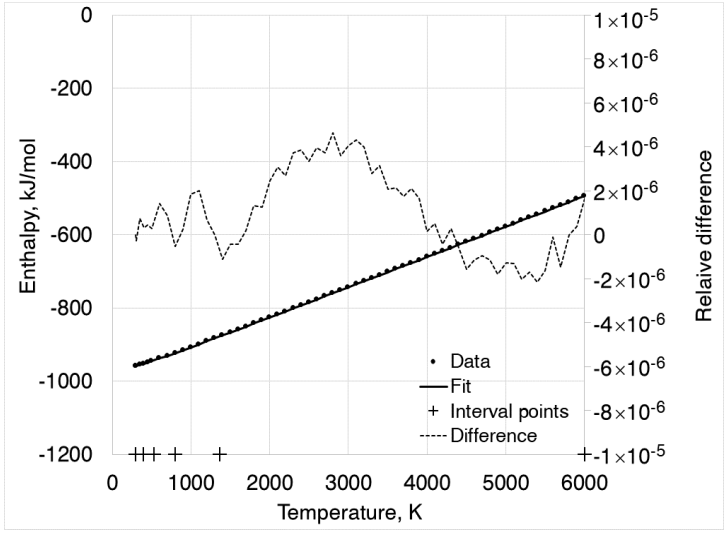


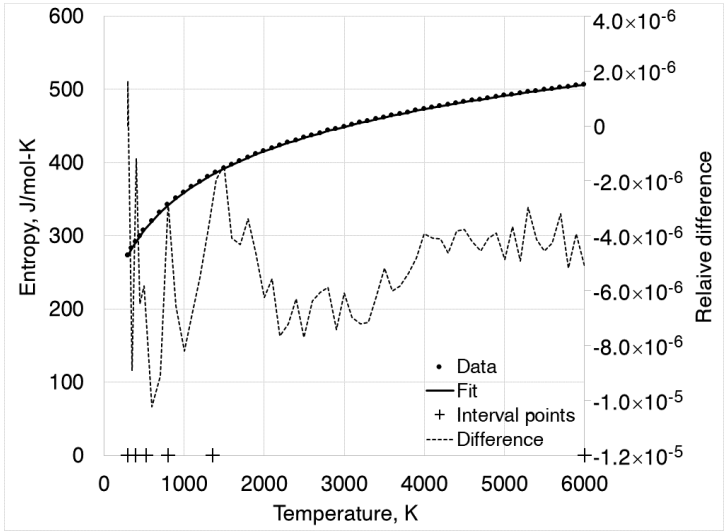


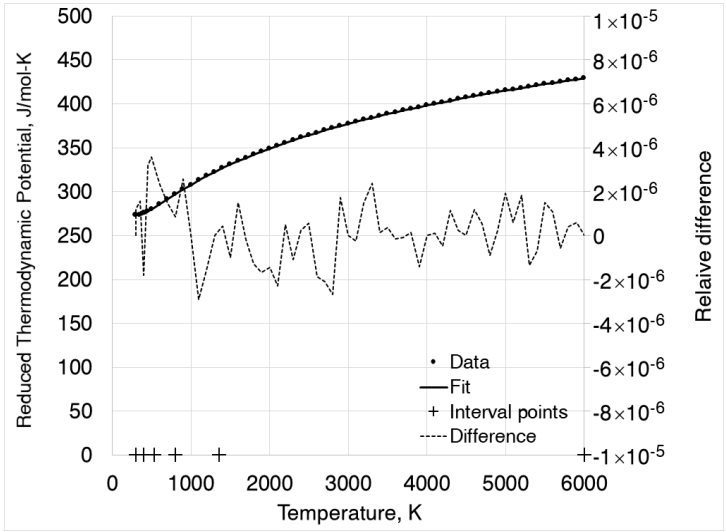
Phosphorus Fluoride (PF_3) – F3P1(g)

MW = 87.969 g/mol – $\Delta_f H^o$(298.15 K) = -958.441 kJ/mol – S^o(298.15 K) = 273.064 J/mol-K – p^o = 0.1 MPa



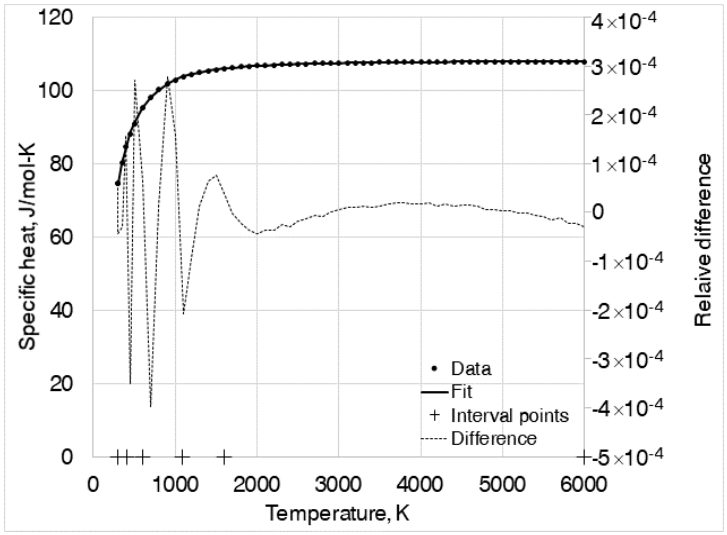


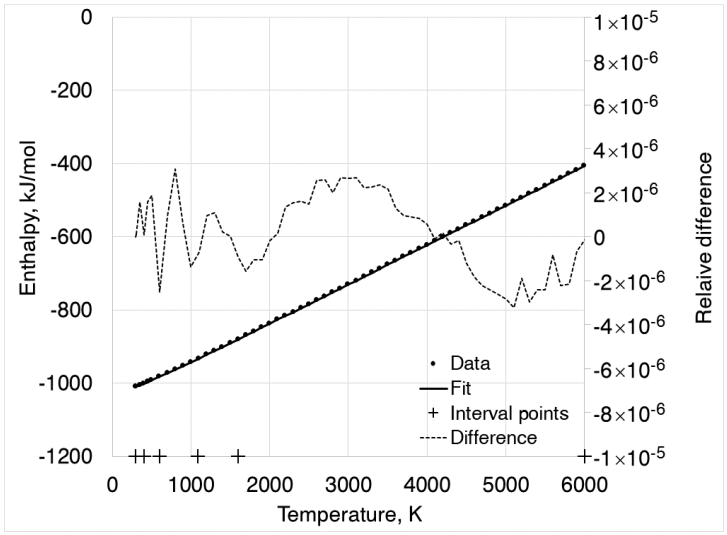


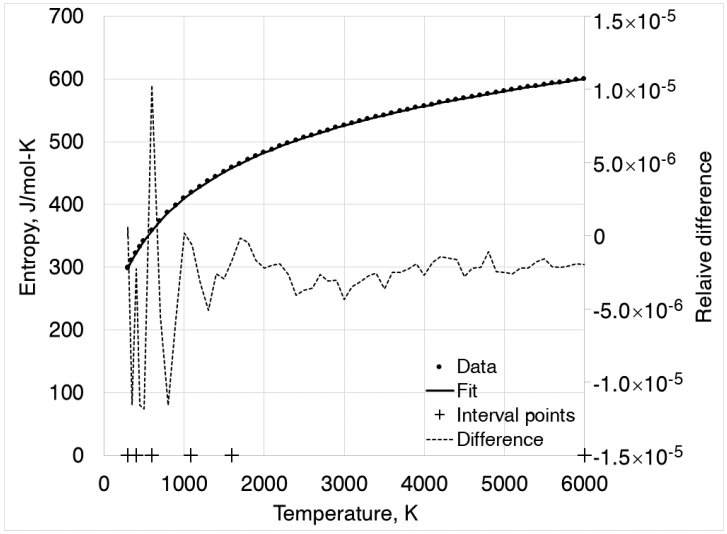


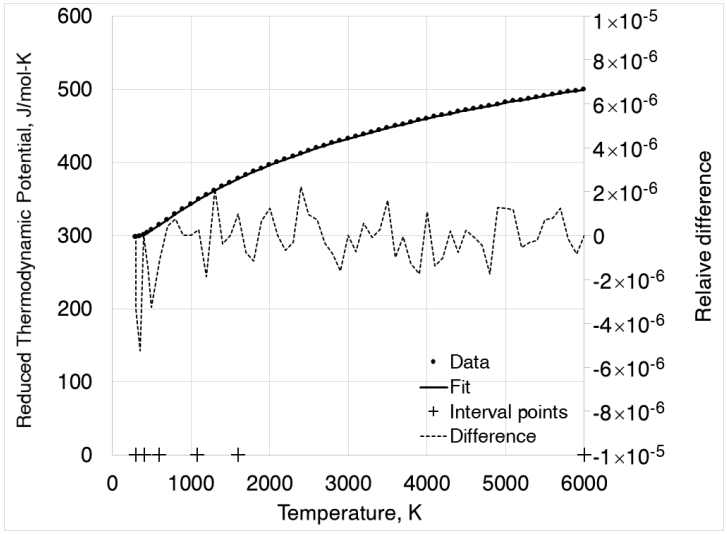
Thiophosphoryl Fluoride (PSF_3) – F3P1S1(g)

MW = 120.03 g/mol – $\Delta_f H^o$(298.15 K) = -1009.068 kJ/mol – S^o(298.15 K) = 298.139 J/mol-K – p^o = 0.1 MPa



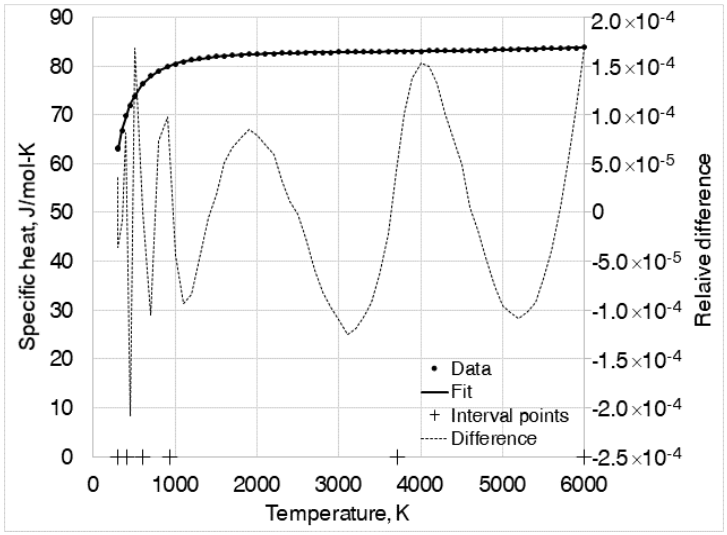


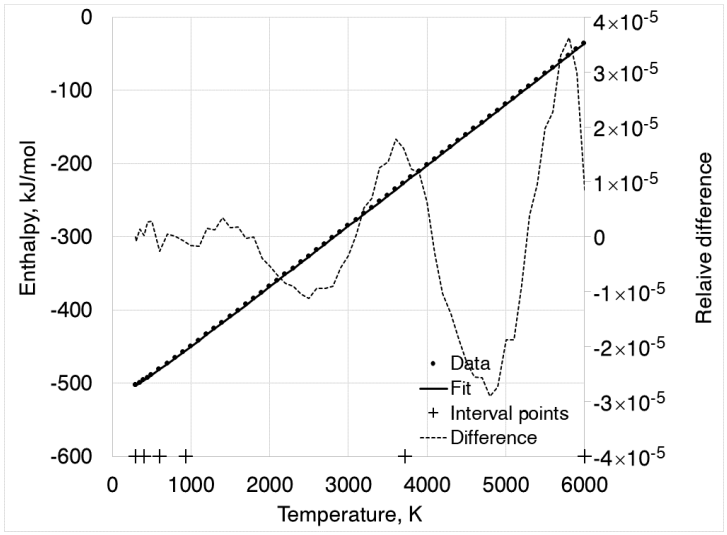


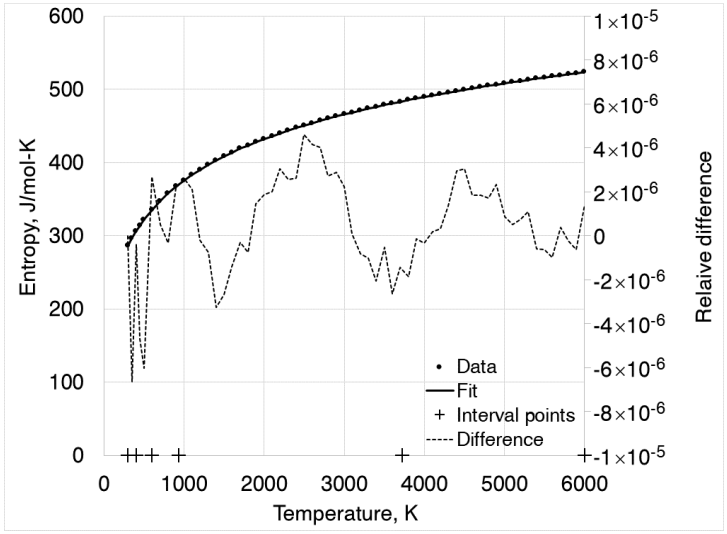


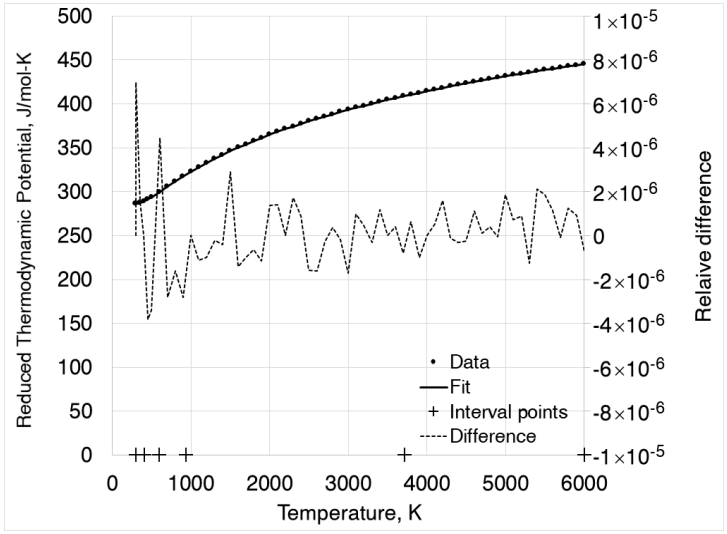
Sulfur Fluoride (SF_3) – F3S1(g)

MW = 89.06 g/mol – $\Delta_f H^o$(298.15 K) = -503.03 kJ/mol – S^o(298.15 K) = 286.174 J/mol-K – p^o = 0.1 MPa



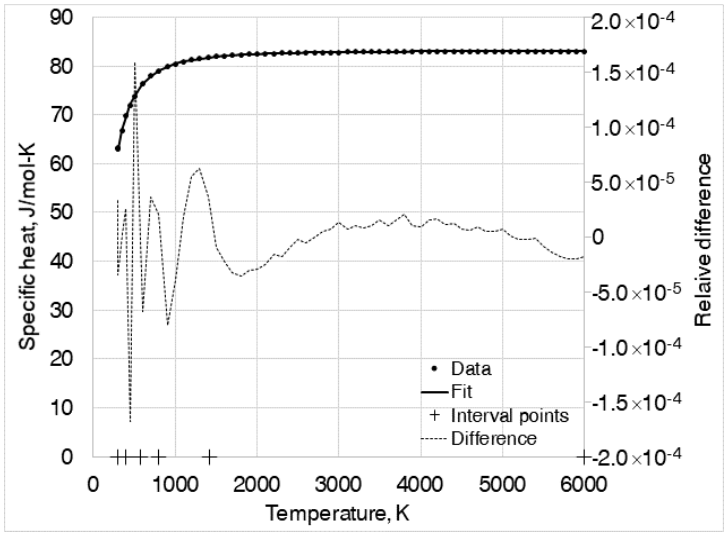


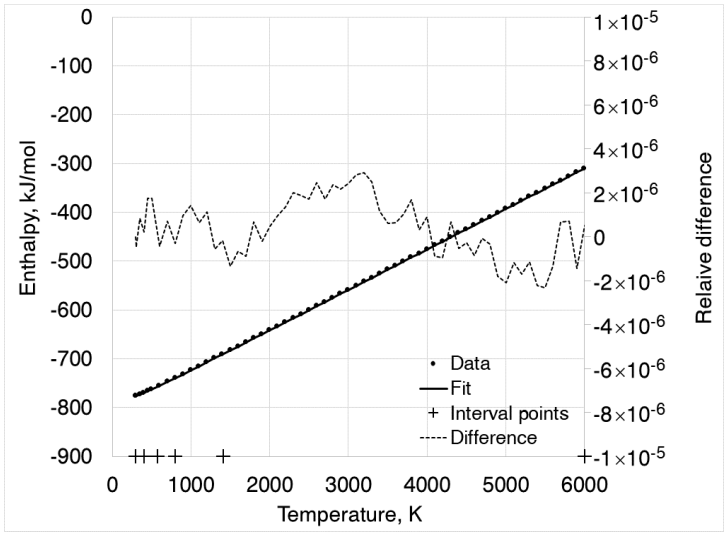


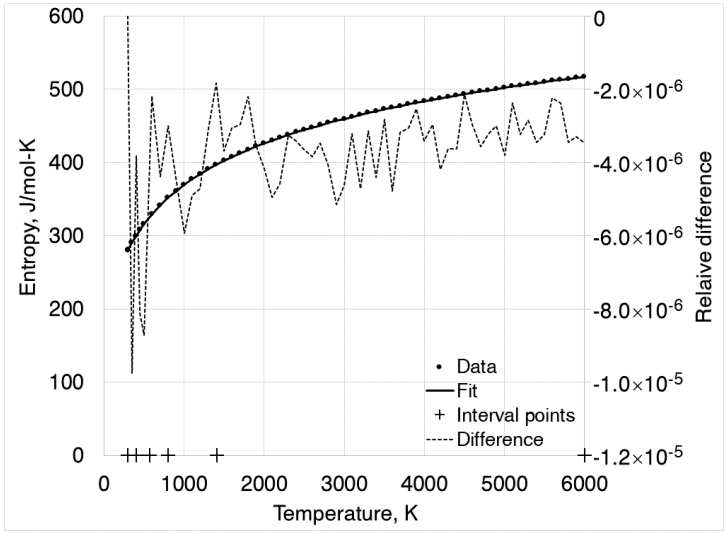


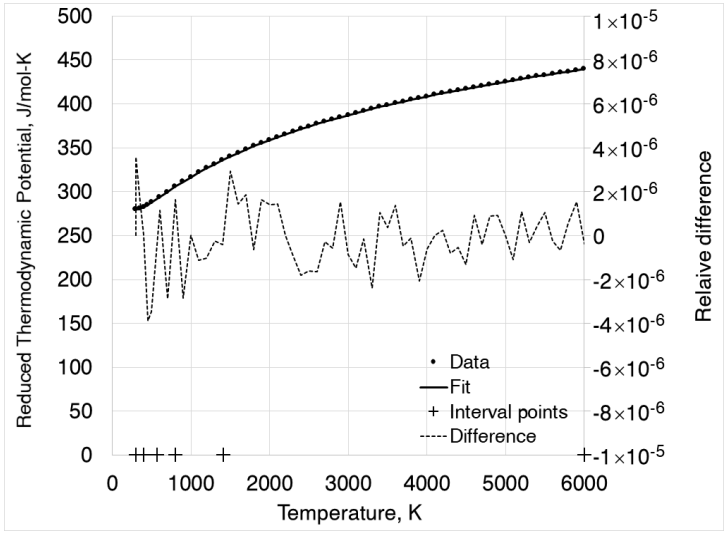
Sulfur Fluoride, Ion (SF_3^-) – F3S1-(g)

MW = 89.06 g/mol – $\Delta_f H^o$(298.15 K) = -776.396 kJ/mol – S^o(298.15 K) = 280.198 J/mol-K – p^o = 0.1 MPa



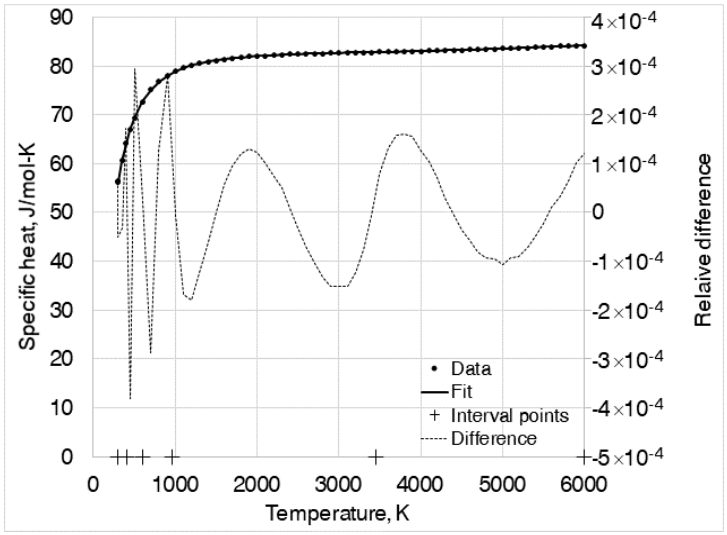


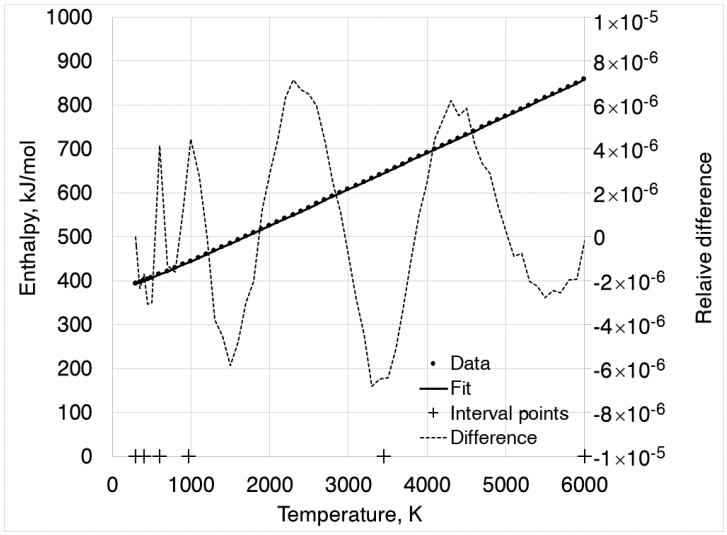


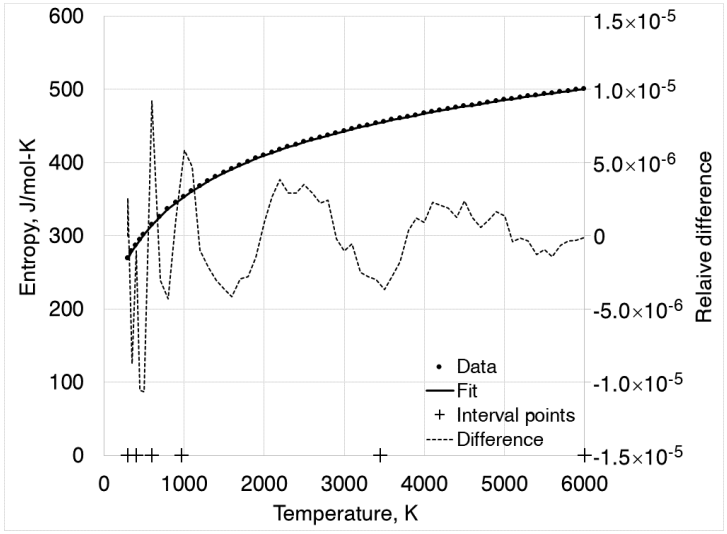


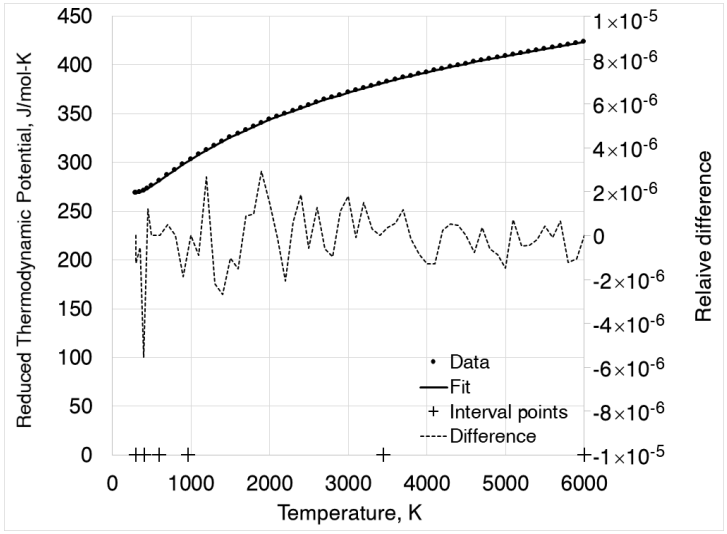
Sulfur Fluoride, Ion (SF_3^+) – F3S1+(g)

MW = 89.05 g/mol – $\Delta_f H^o$(298.15 K) = 393.589 kJ/mol – S^o(298.15 K) = 268.972 J/mol-K – p^o = 0.1 MPa



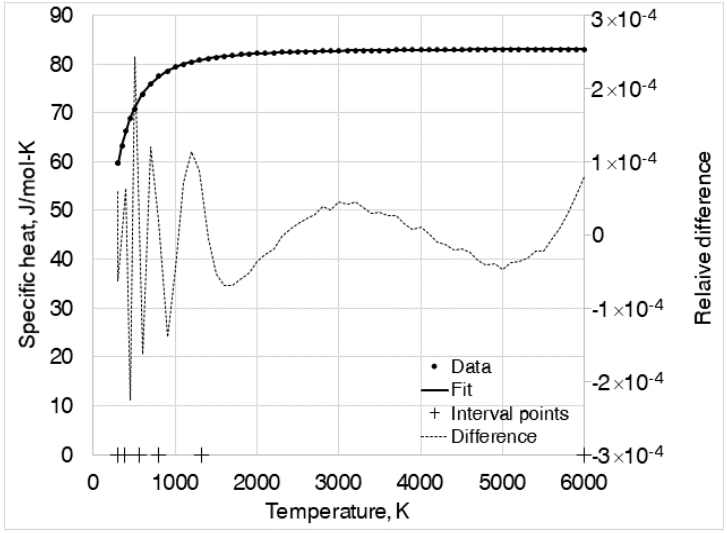


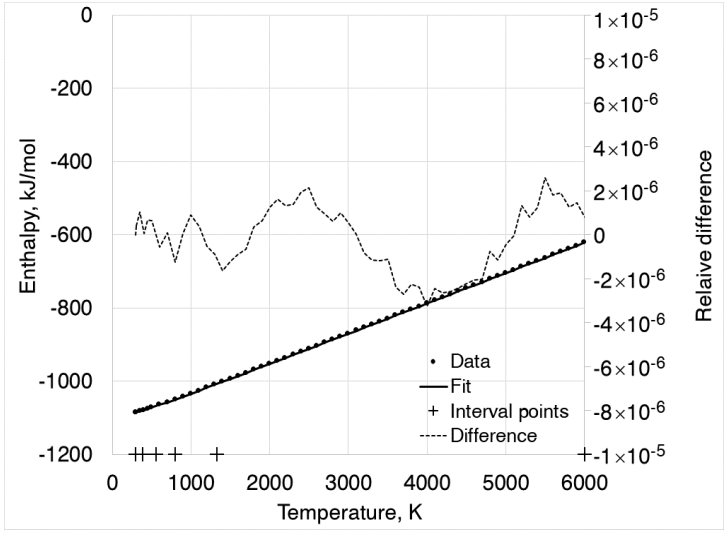


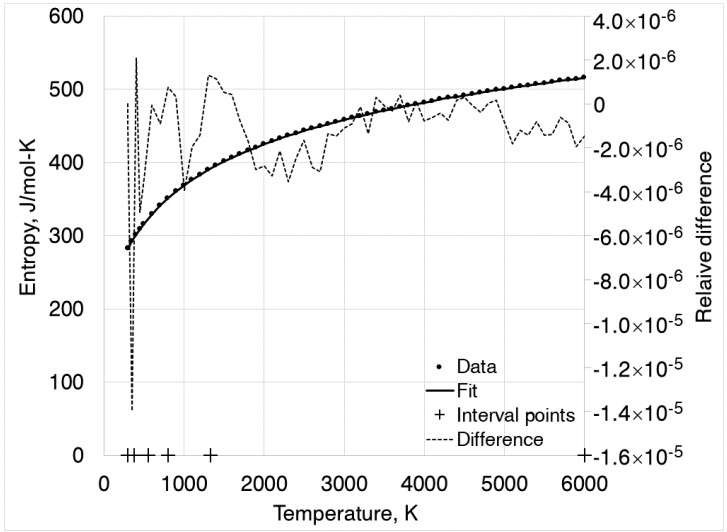


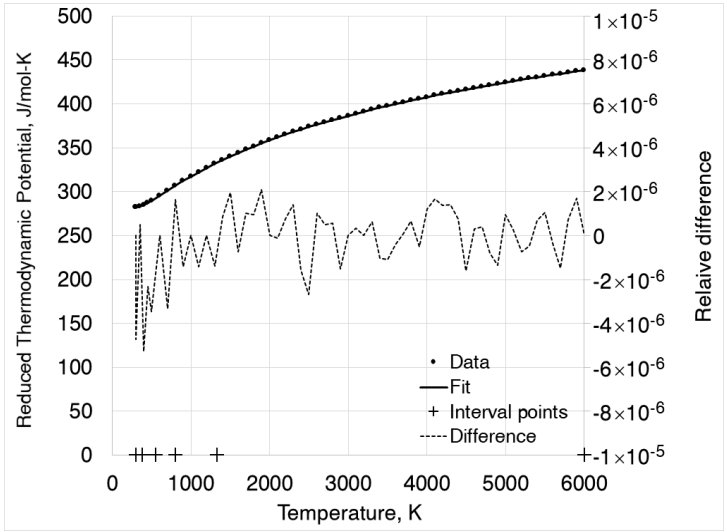
Trifluorosilyl (SiF_3) – F3Si1(g)

MW = 85.0807 g/mol – $\Delta_f H^o$(298.15 K) = -1085.33 kJ/mol – S^o(298.15 K) = 282.381 J/mol-K – p^o = 0.1 MPa



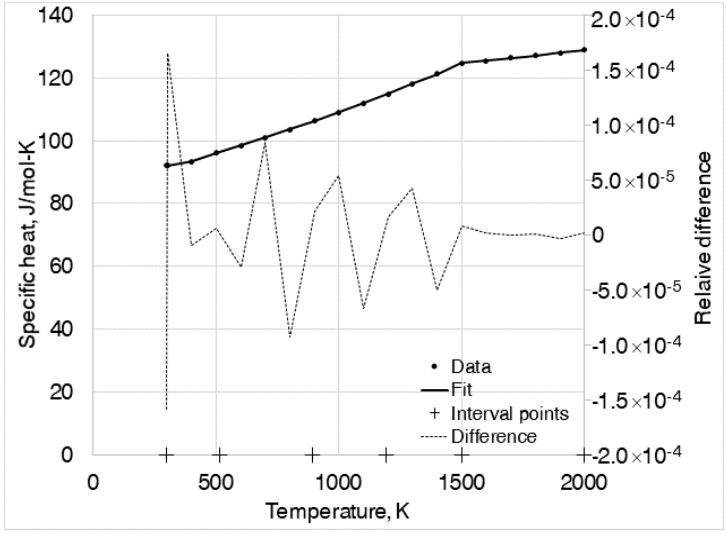


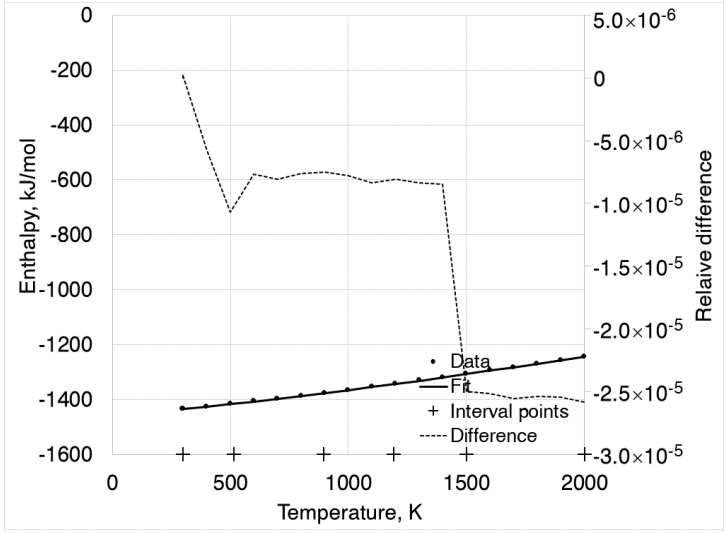


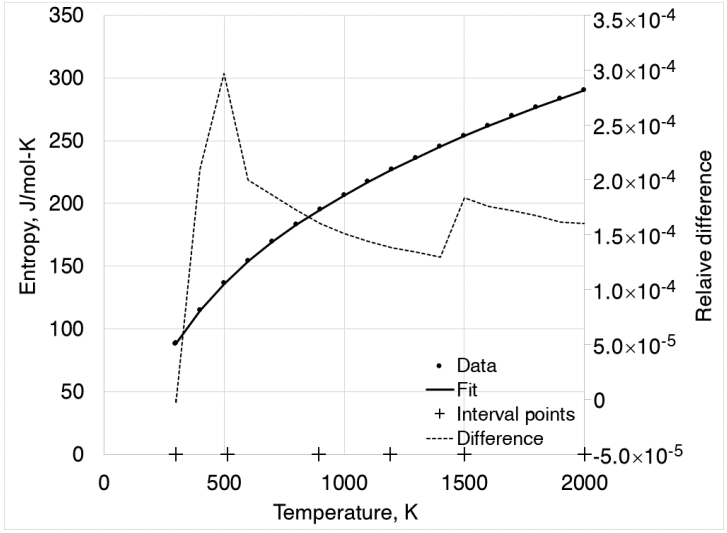


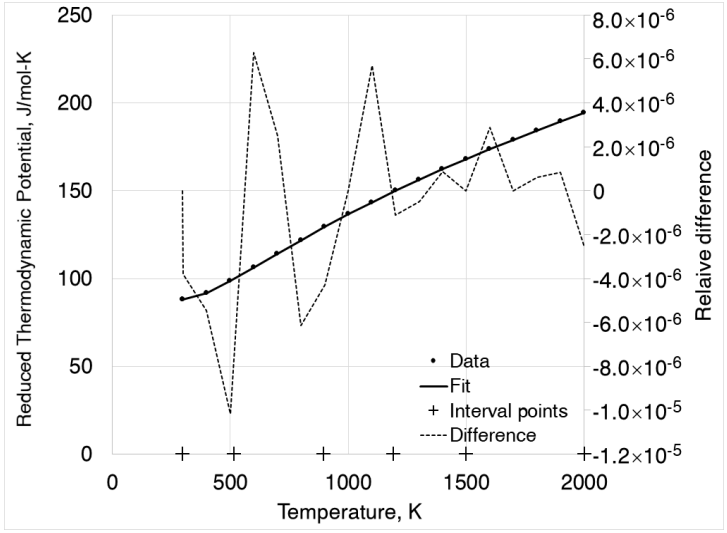
Titanium Fluoride (TiF_3) – F3Ti1(cr)

MW = 104.88 g/mol – $\Delta_f H^o$(298.15 K) = -1435.53 kJ/mol – S^o(298.15 K) = 87.864 J/mol-K – p^o = 0.1 MPa



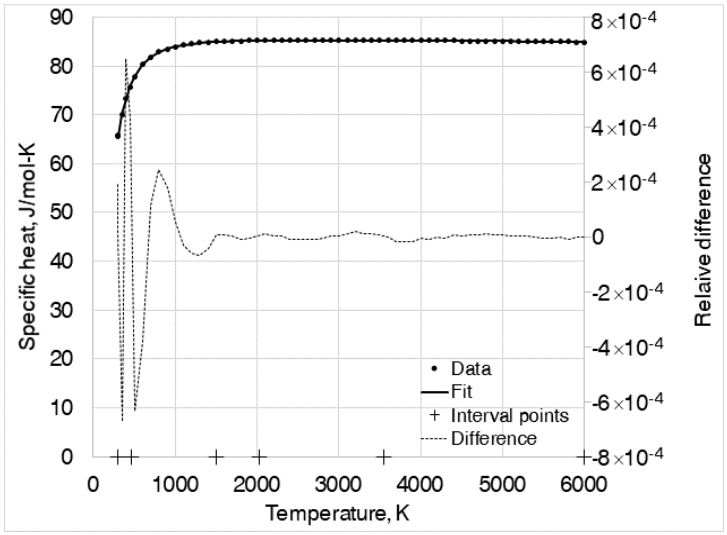


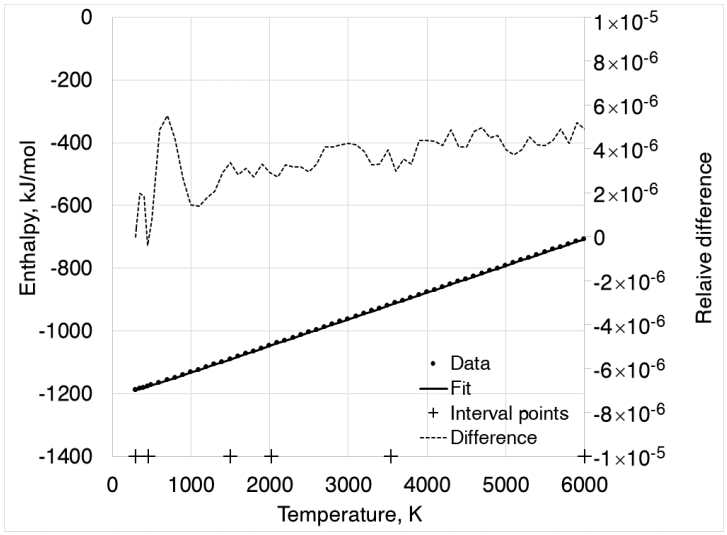


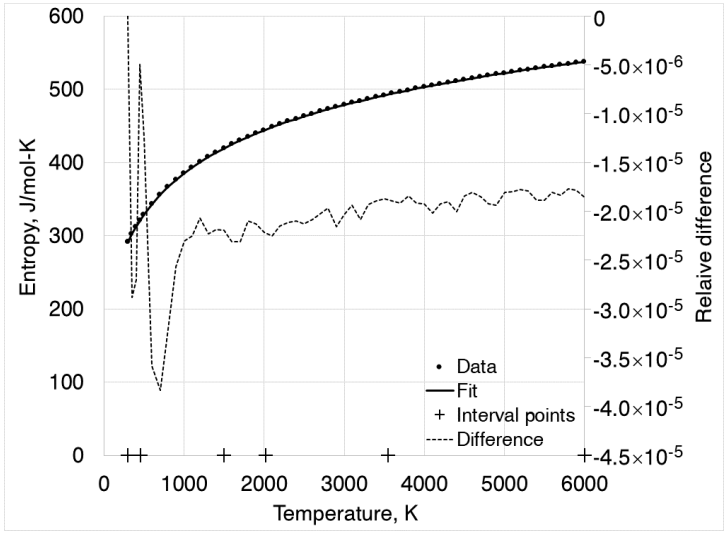


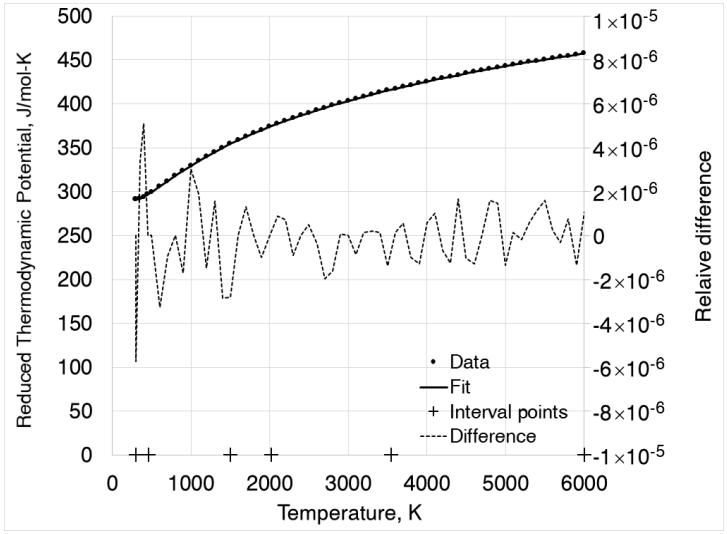
Titanium Fluoride (TiF_3) – F3Ti1(g)

MW = 104.88 g/mol – $\Delta_f H^o$(298.15 K) = -1188.674 kJ/mol – S^o(298.15 K) = 291.44 J/mol-K – p^o = 0.1 MPa



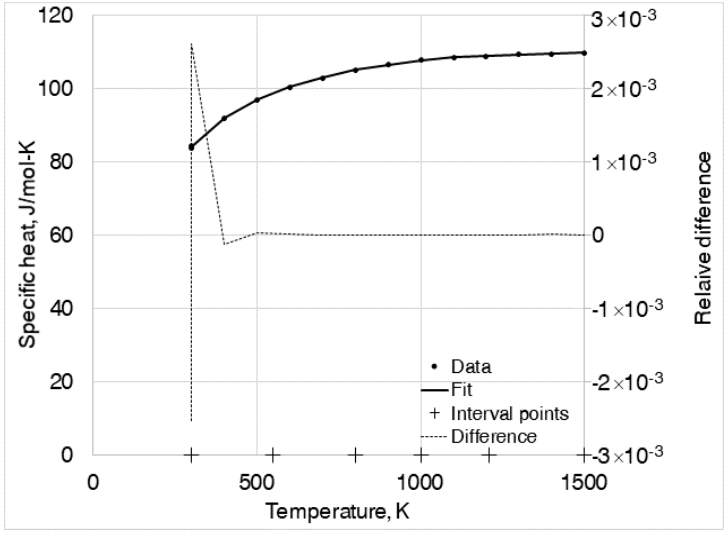


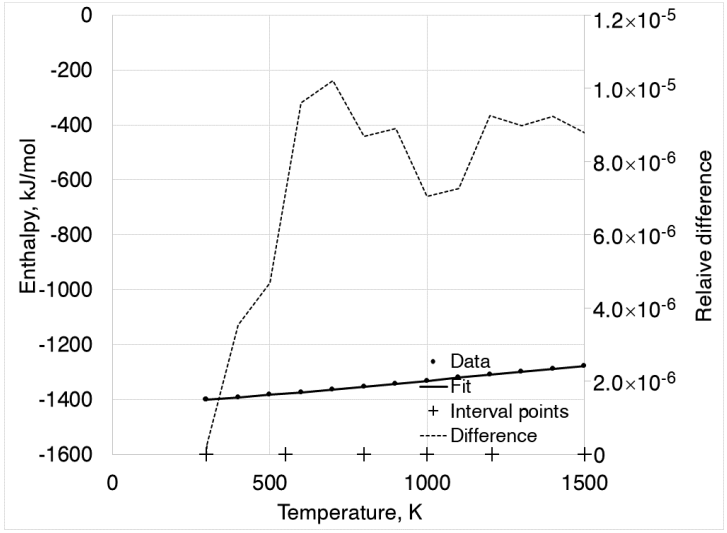


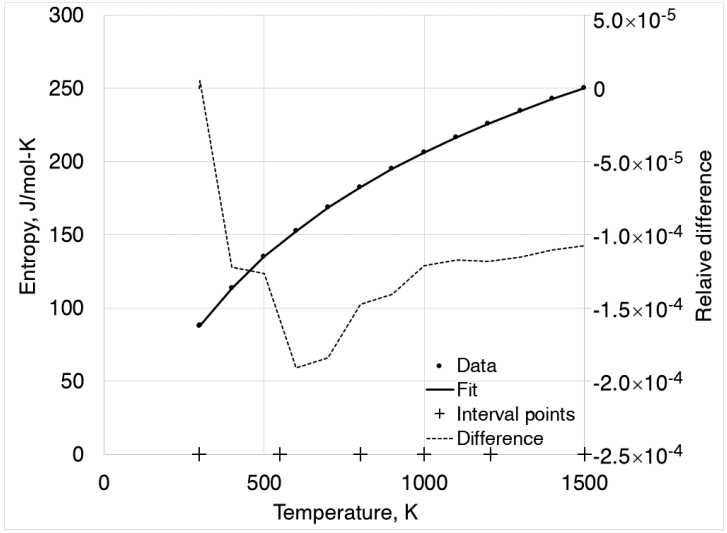


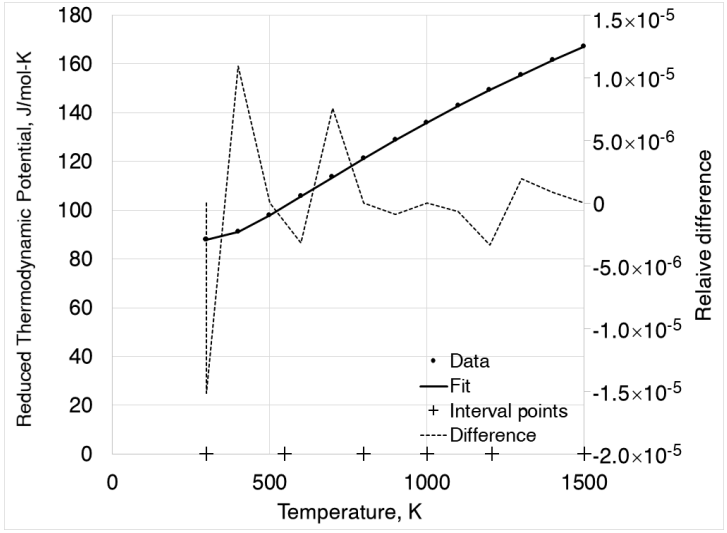
Zirconium Fluoride (ZrF_3) – F3Zr1(cr)

MW = 148.22 g/mol – $\Delta_f H^o$(298.15 K) = -1401.64 kJ/mol – S^o(298.15 K) = 87.864 J/mol-K – p^o = 0.1 MPa



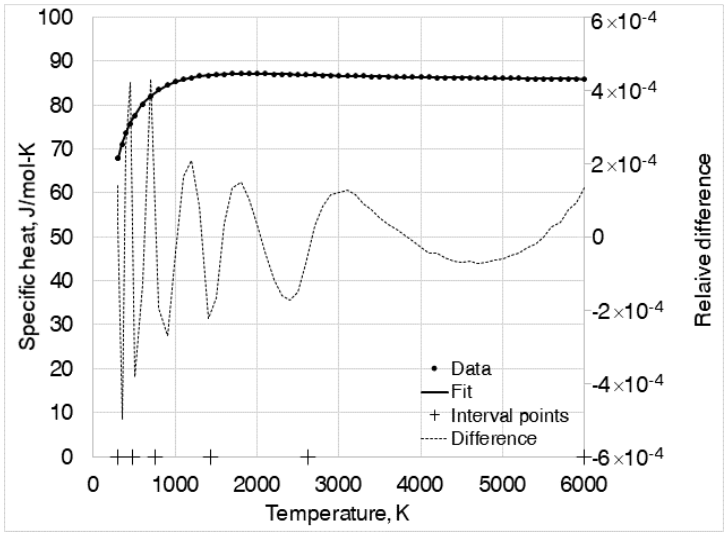


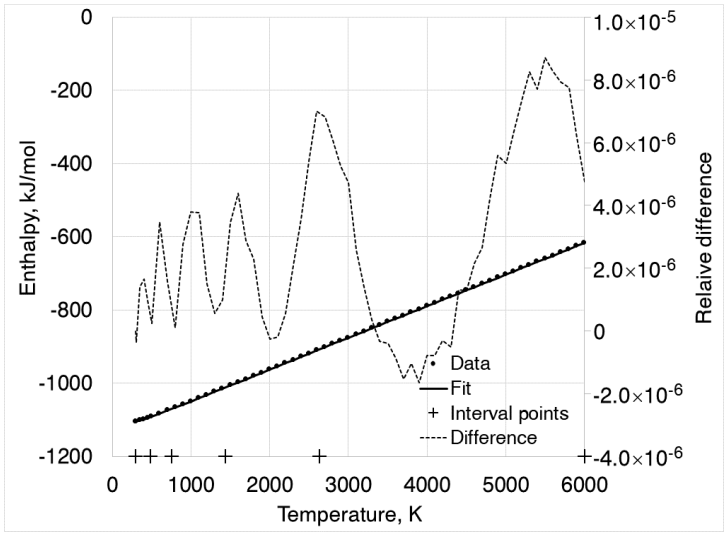


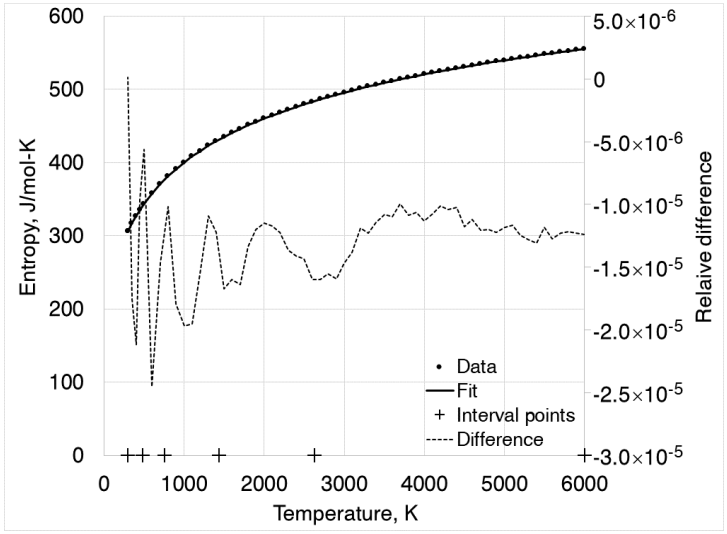


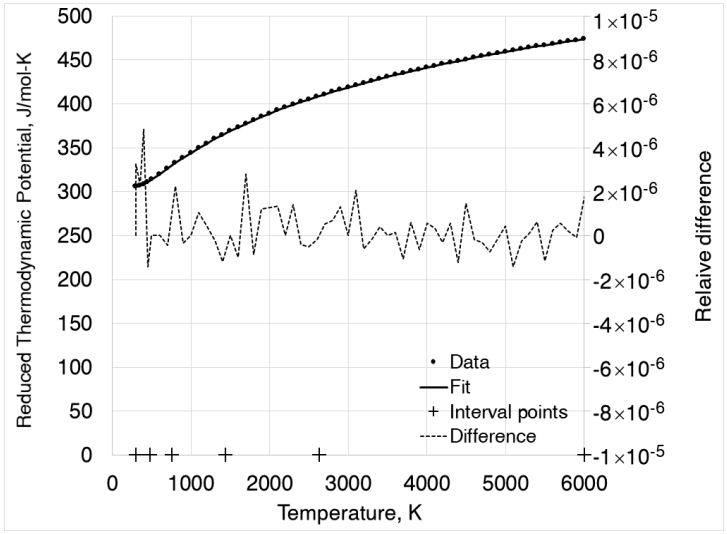
Zirconium Fluoride (ZrF_3) – F3Zr1(g)

MW = 148.22 g/mol – $\Delta_f H^o$(298.15 K) = -1105.413 kJ/mol – S^o(298.15 K) = 305.742 J/mol-K – p^o = 0.1 MPa



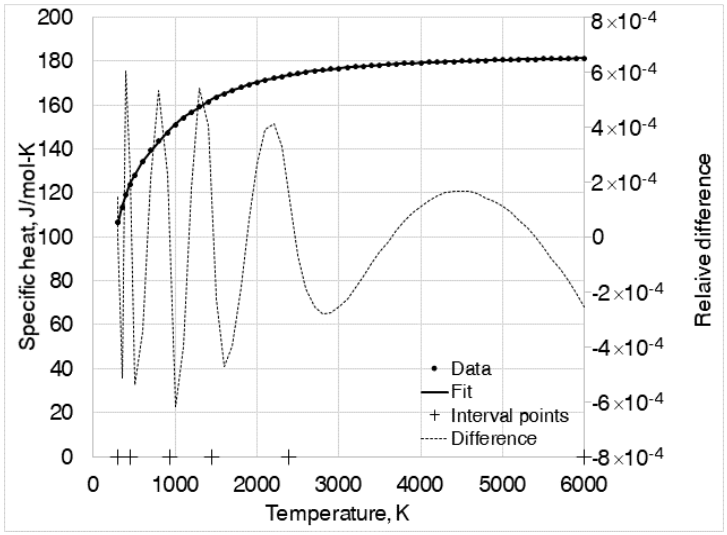


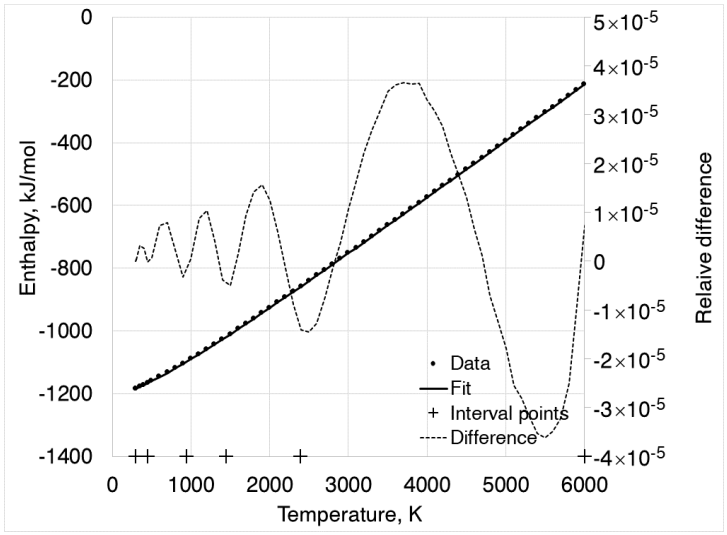


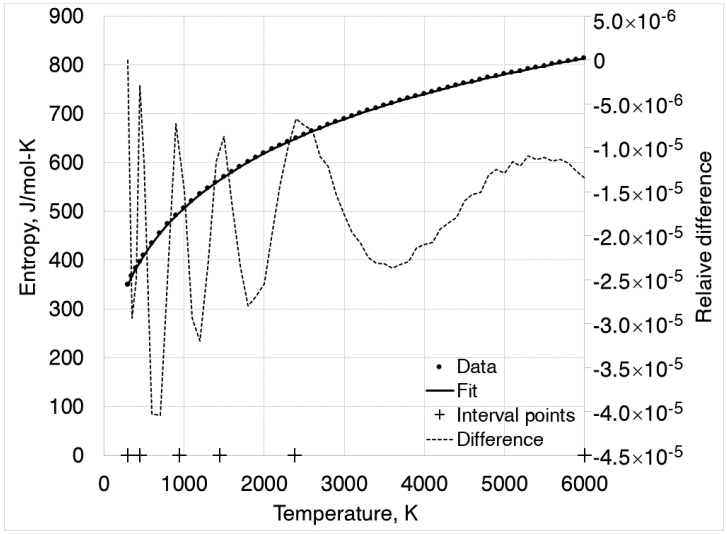


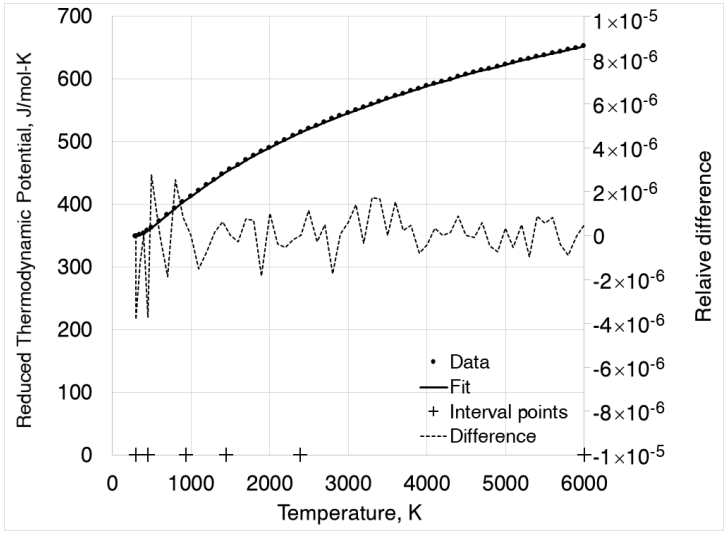
Hydrogen Fluoride (H_4F_4) – F4H4(g)

MW = 80.0254 g/mol – $\Delta_f H^o$(298.15 K) = -1183.821 kJ/mol – S^o(298.15 K) = 349.21 J/mol-K – p^o = 0.1 MPa



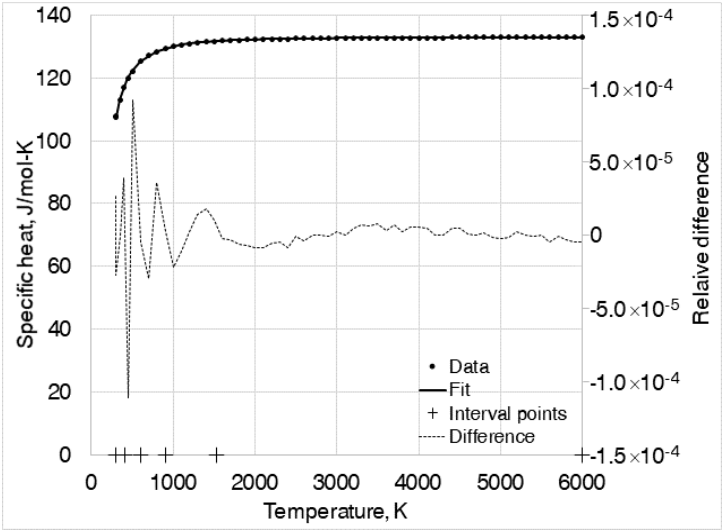


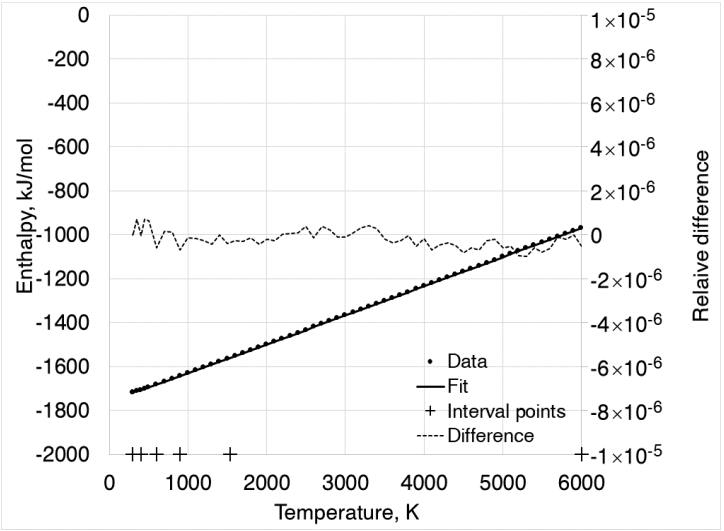


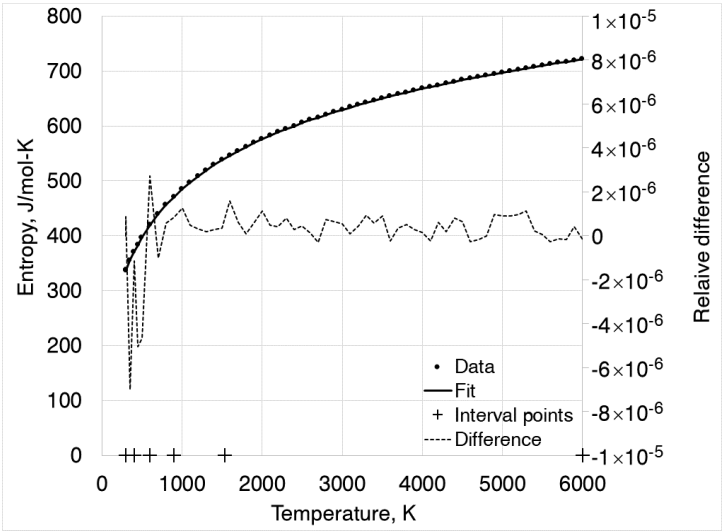


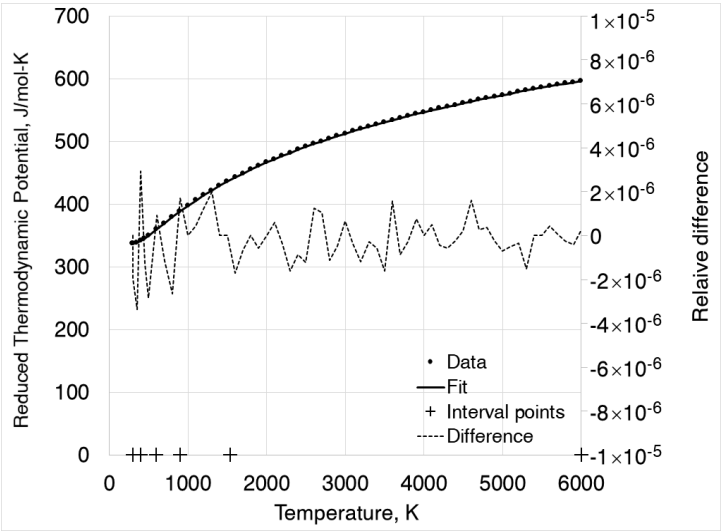
Magnesium Fluoride ($(MgF_2)_2$) – F4Mg2(g)

MW = 124.604 g/mol – $\Delta_f H^o$(298.15 K) = -1718.369 kJ/mol – S^o(298.15 K) = 337.014 J/mol-K – p^o = 0.1 MPa



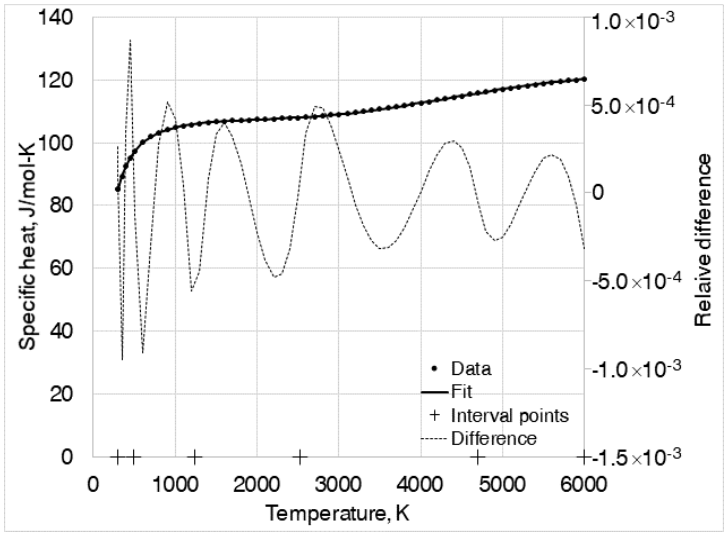


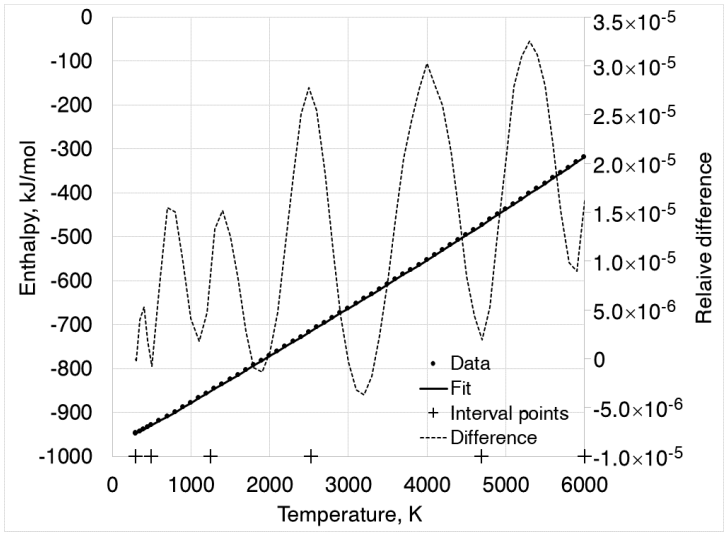


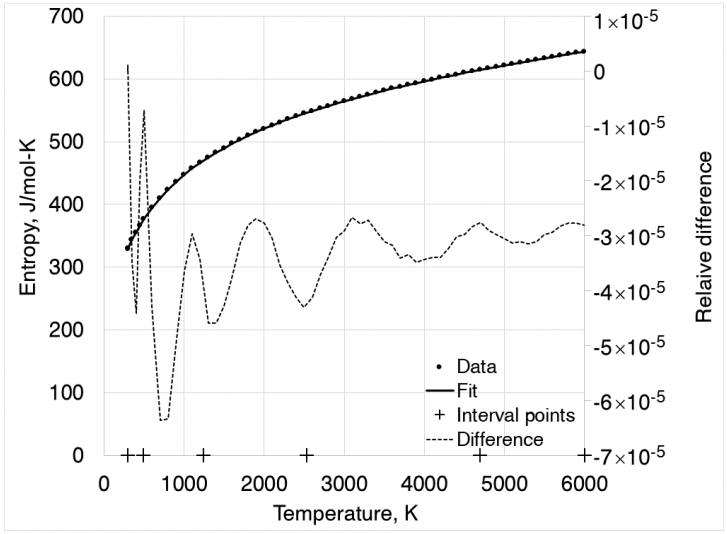


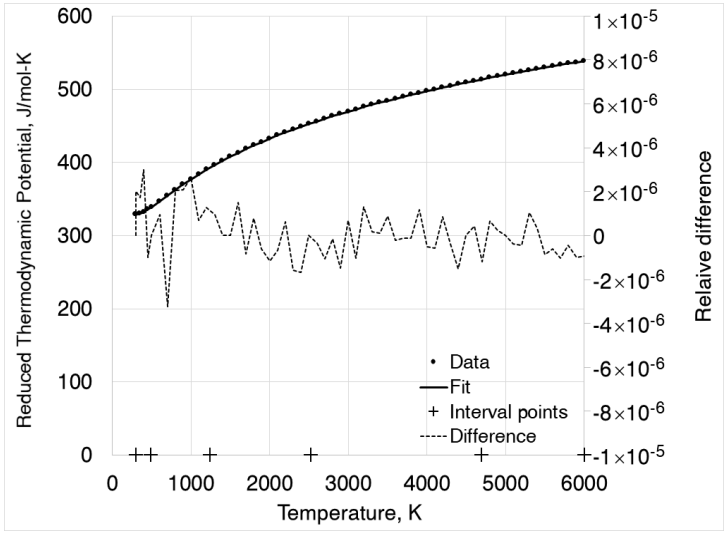
Molybdenum Fluoride (MoF_4) – F4Mo1(g)

MW = 171.93 g/mol – $\Delta_f H^o$(298.15 K) = -947.676 kJ/mol – S^o(298.15 K) = 328.94 J/mol-K – p^o = 0.1 MPa



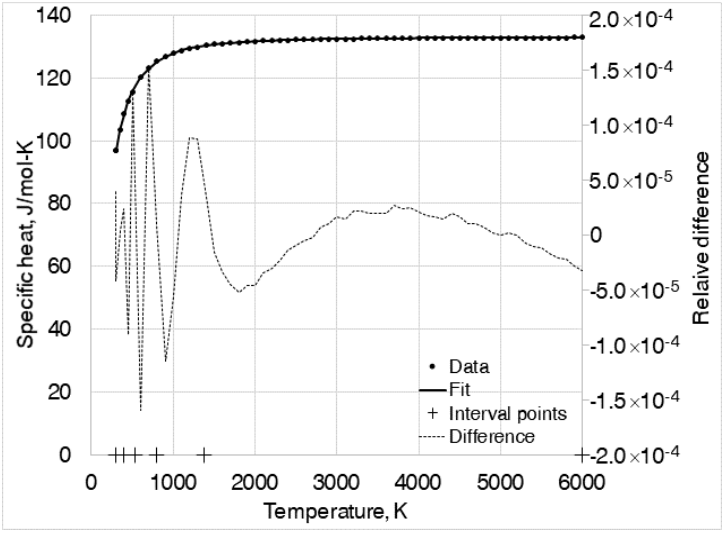


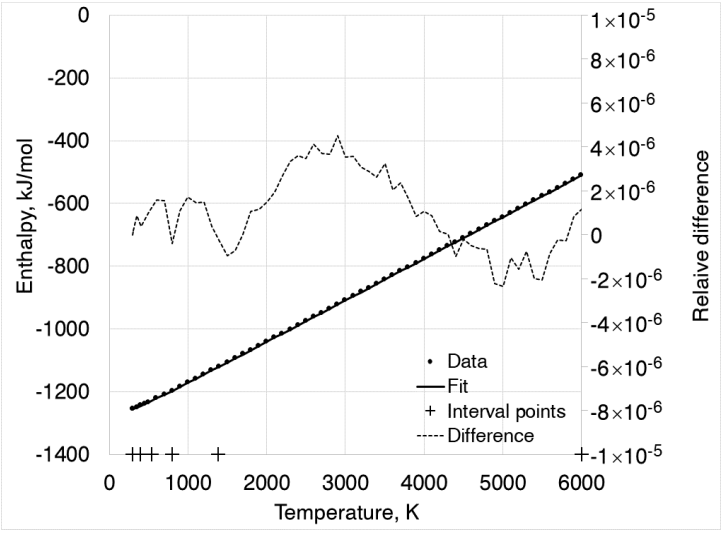


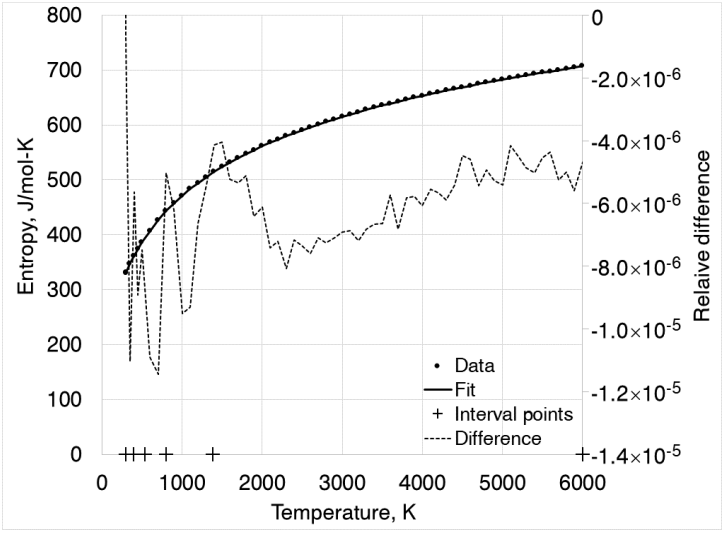


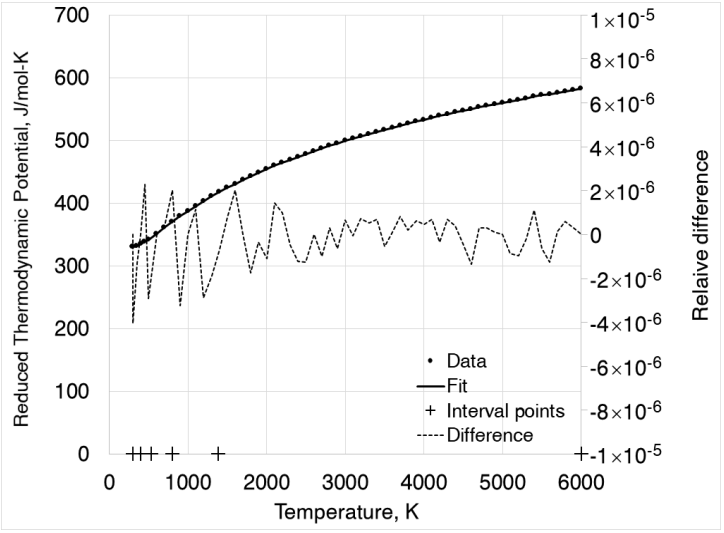
Molybdenum Fluoride Oxide (MoF_4O) – F4Mo1O1(g)

MW = 187.93 g/mol – $\Delta_f H^o$(298.15 K) = -1255.2 kJ/mol – S^o(298.15 K) = 330.617 J/mol-K – p^o = 0.1 MPa



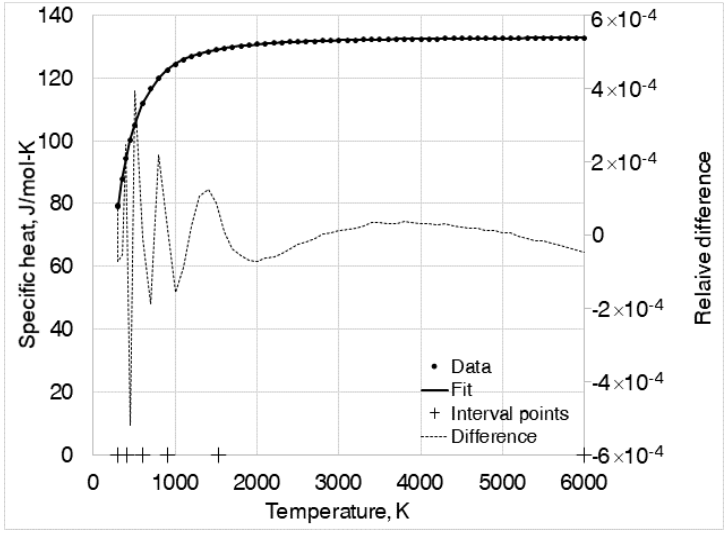


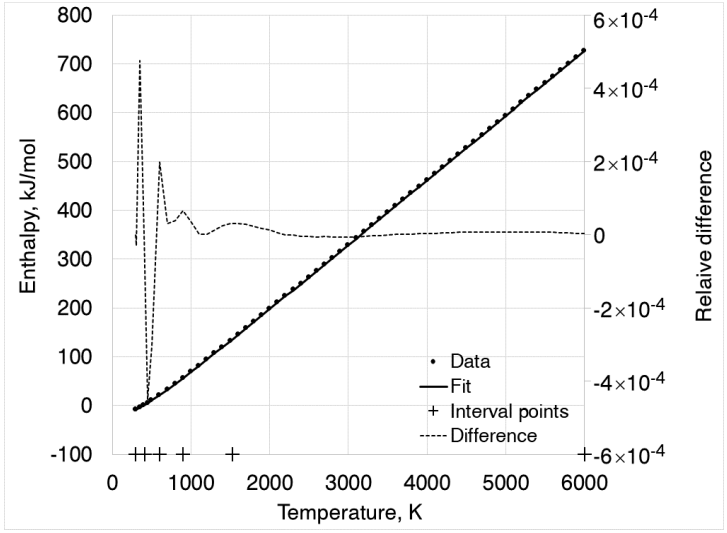


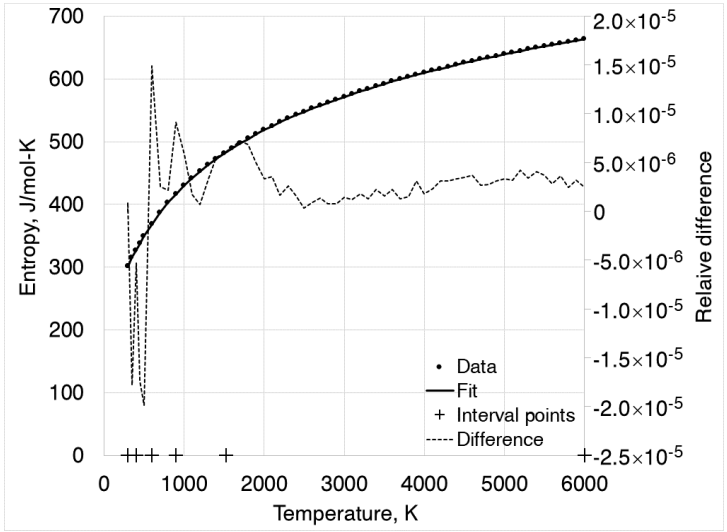


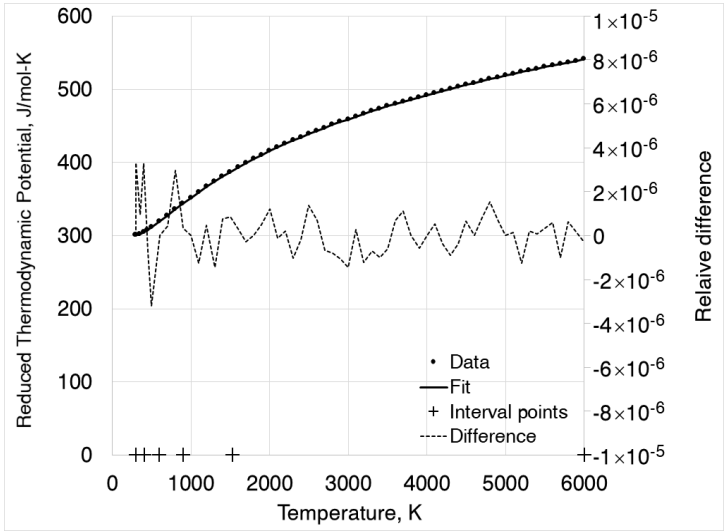
Tetrafluorohydrazine (N_2F_4) – F4N2(g)

MW = 104.007 g/mol – $\Delta_f H^o$(298.15 K) = -8.368 kJ/mol – S^o(298.15 K) = 301.191 J/mol-K – p^o = 0.1 MPa



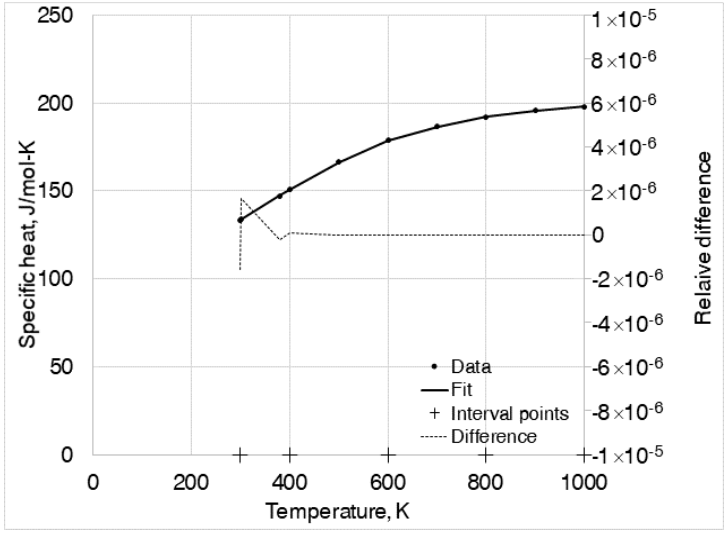


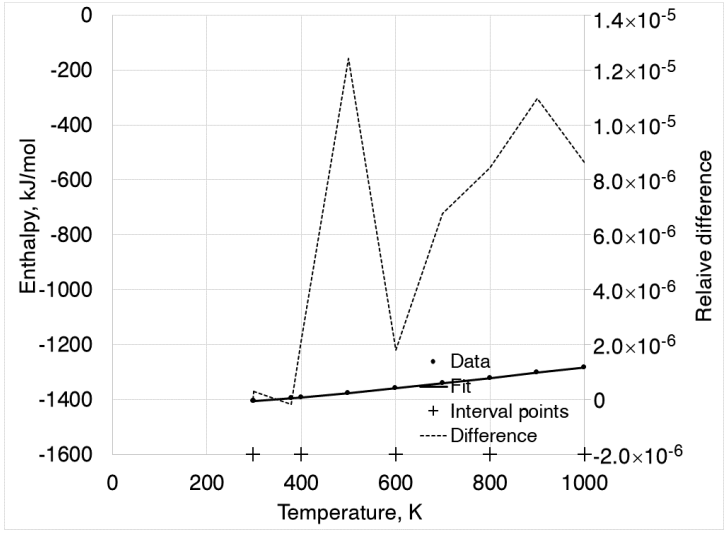


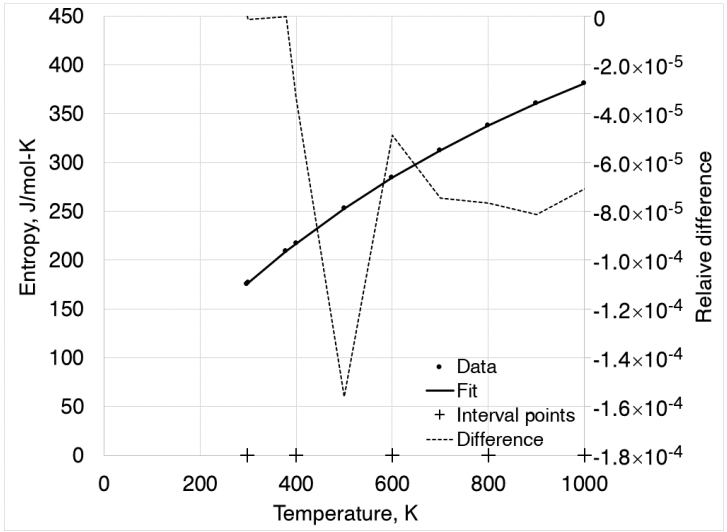


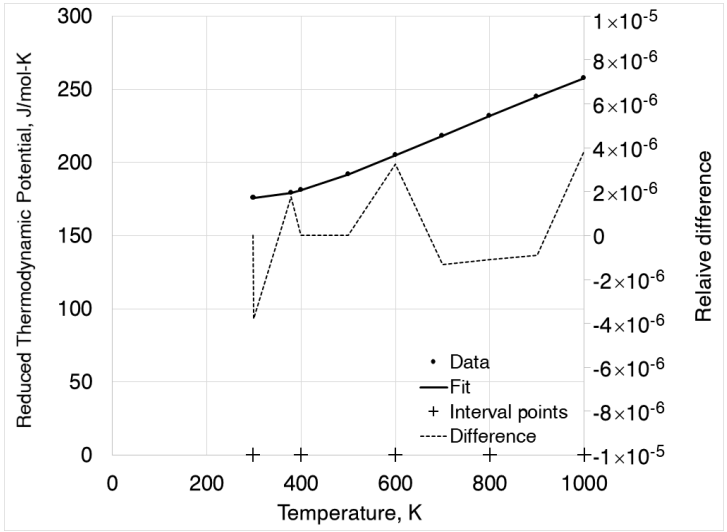
Tungsten Fluoride Oxide (WF_4O) – F4O1W1(cr)

MW = 275.84 g/mol – $\Delta_f H^o$(298.15 K) = -1406.908 kJ/mol – S^o(298.15 K) = 175.728 J/mol-K – p^o = 0.1 MPa



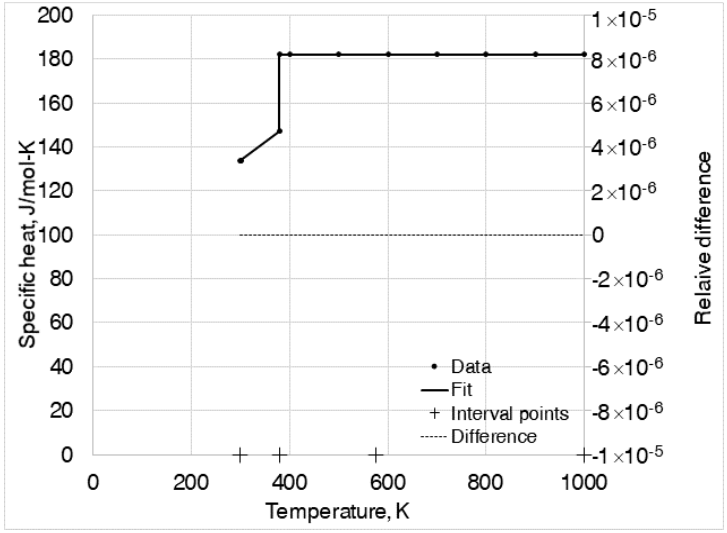


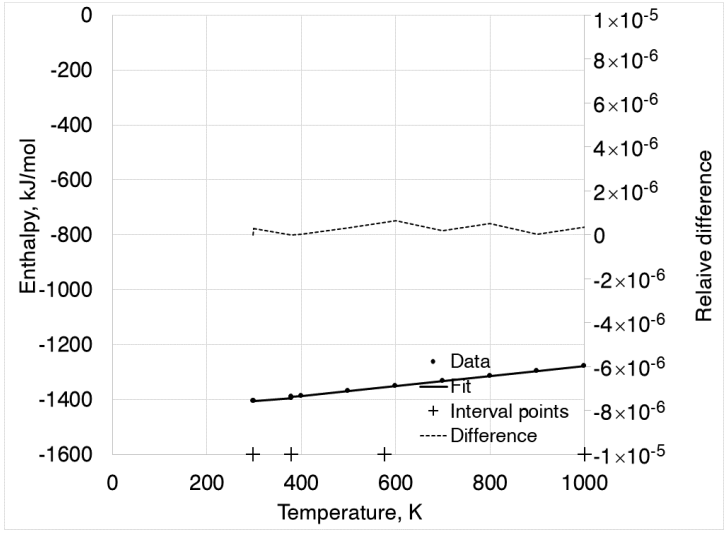


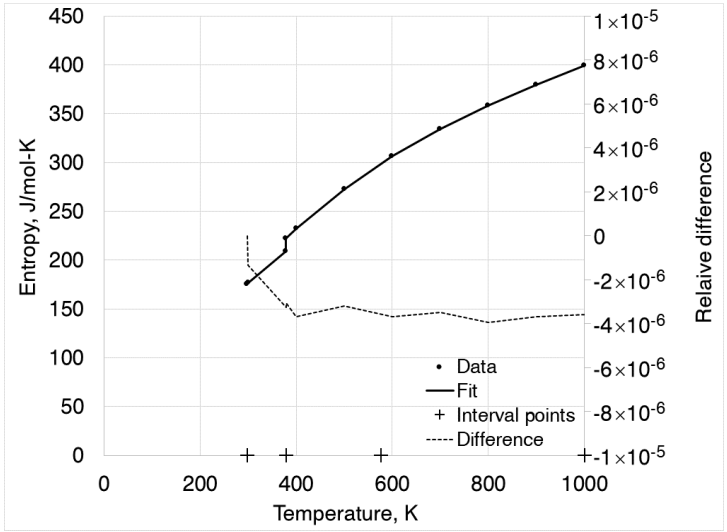


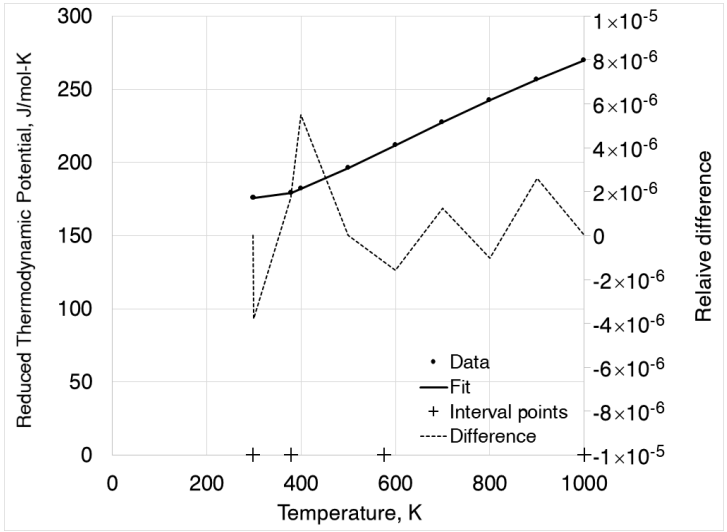
Tungsten Fluoride Oxide (WF_4O) – F4O1W1(cr,l)

MW = 275.84 g/mol – $\Delta_f H^o$(298.15 K) = -1406.908 kJ/mol – S^o(298.15 K) = 175.728 J/mol-K – p^o = 0.1 MPa



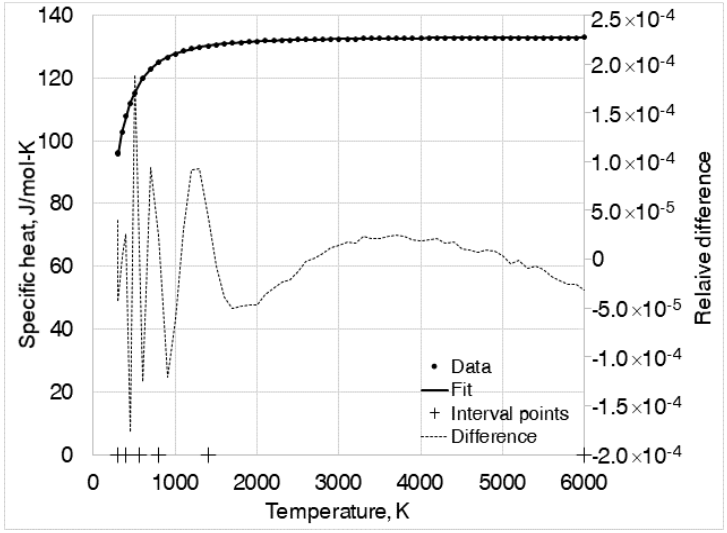


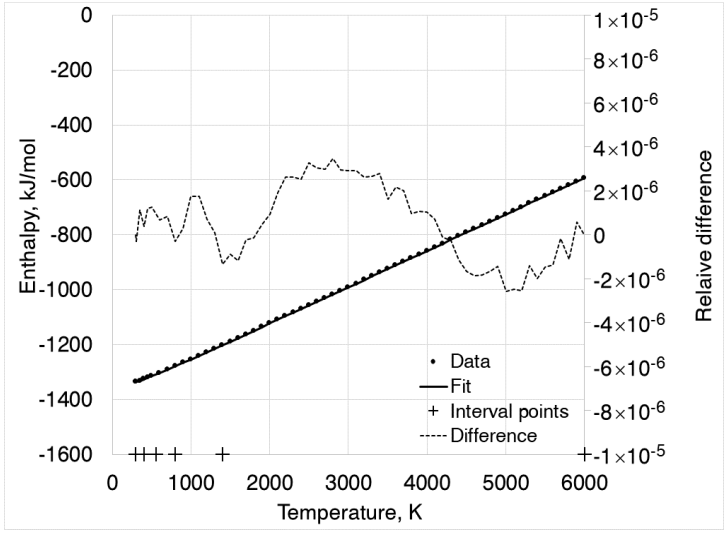


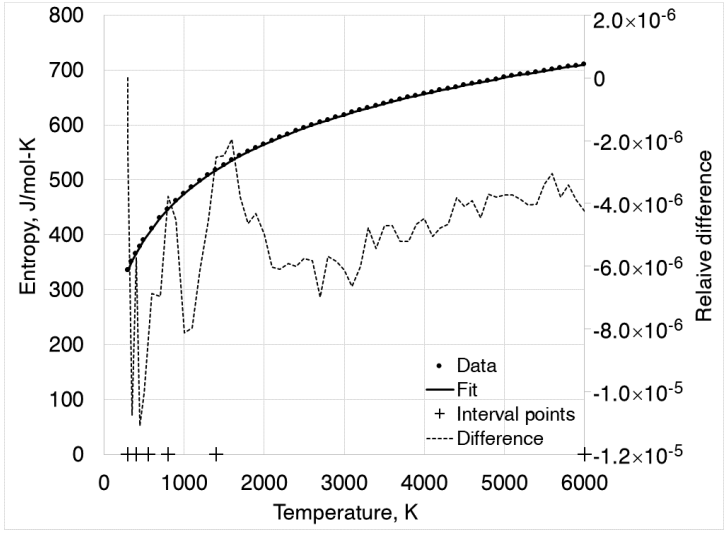


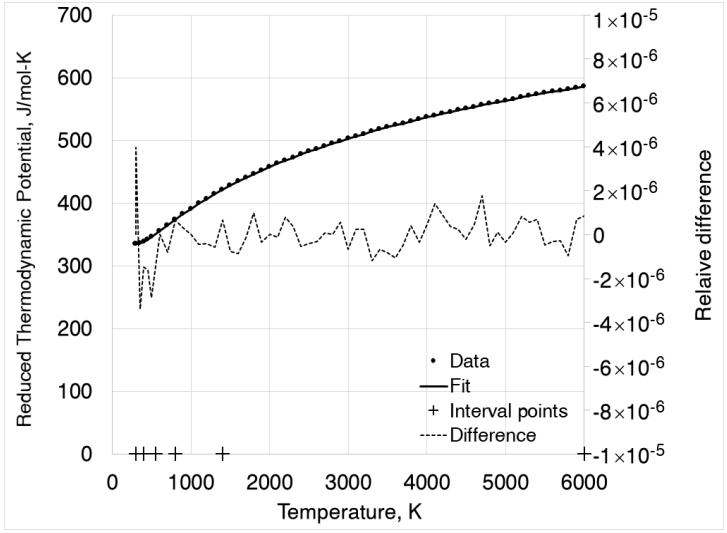
Tungsten Fluoride Oxide (WF_4O) – F4O1W1(g)

MW = 275.84 g/mol – $\Delta_f H^o$(298.15 K) = -1336.621 kJ/mol – S^o(298.15 K) = 334.705 J/mol-K – p^o = 0.1 MPa



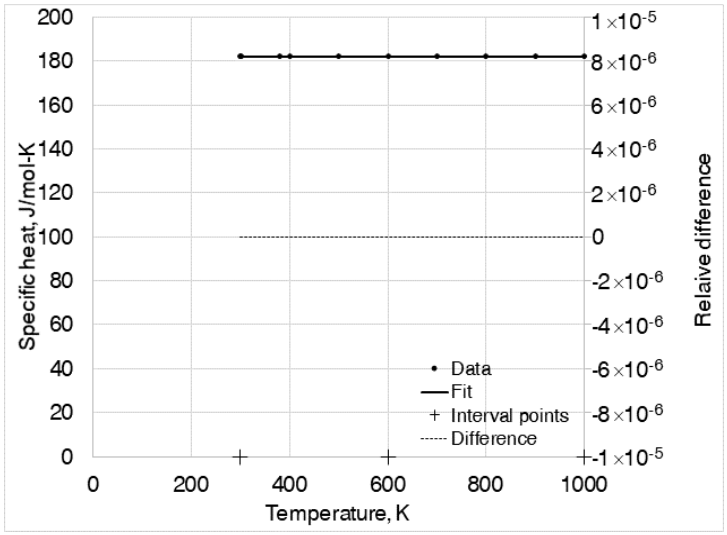


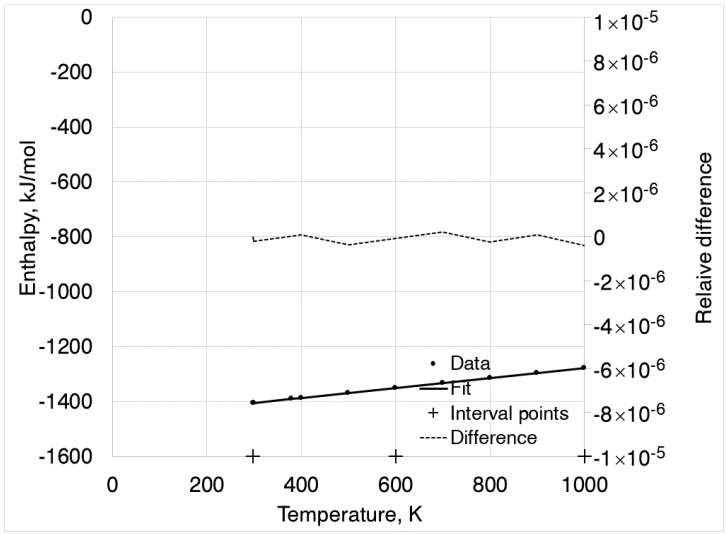


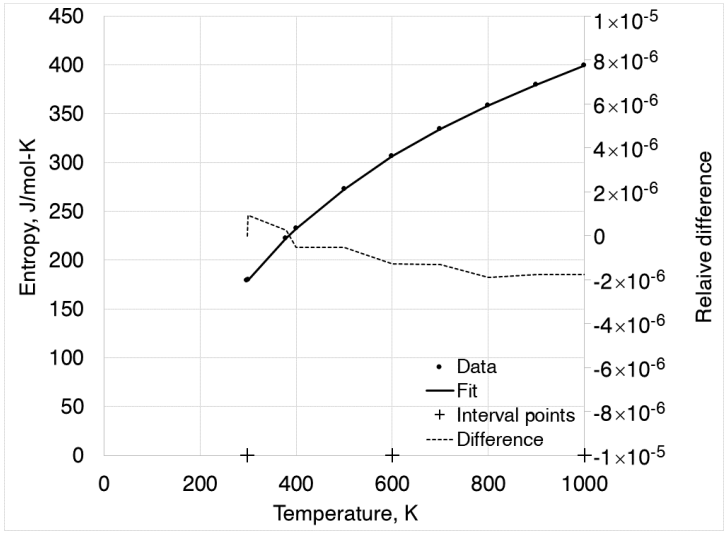


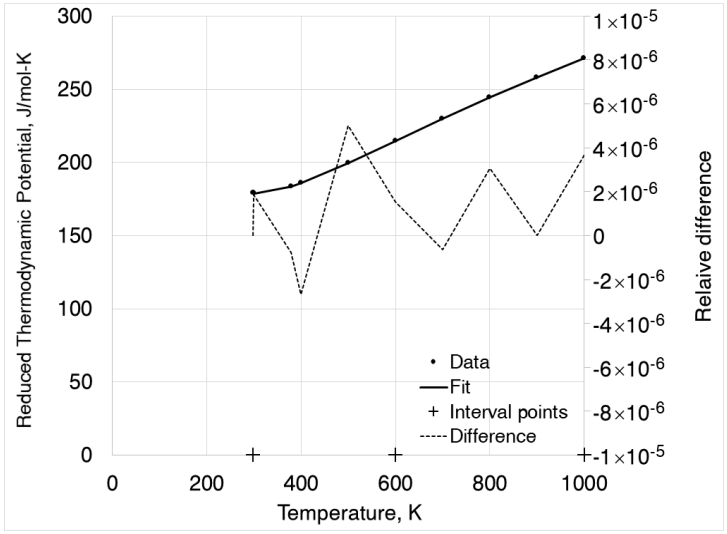
Tungsten Fluoride Oxide (WF_4O) – F4O1W1(l)

MW = 275.84 g/mol – $\Delta_f H^o$(298.15 K) = -1405.255 kJ/mol – S^o(298.15 K) = 178.915 J/mol-K – p^o = 0.1 MPa



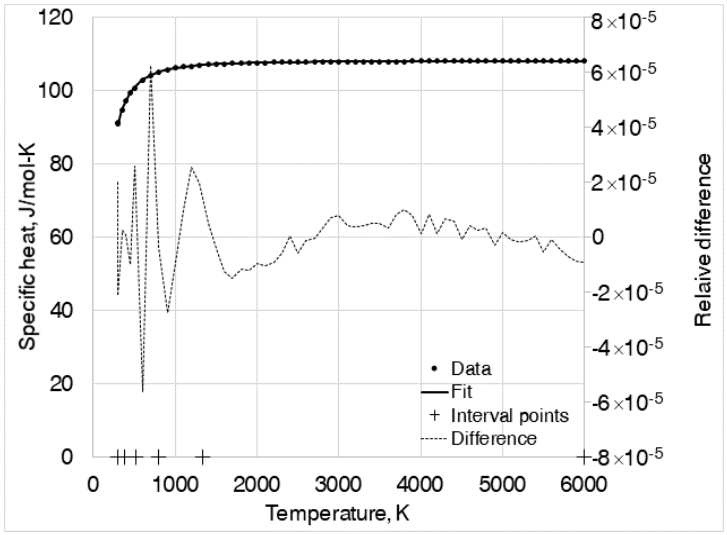


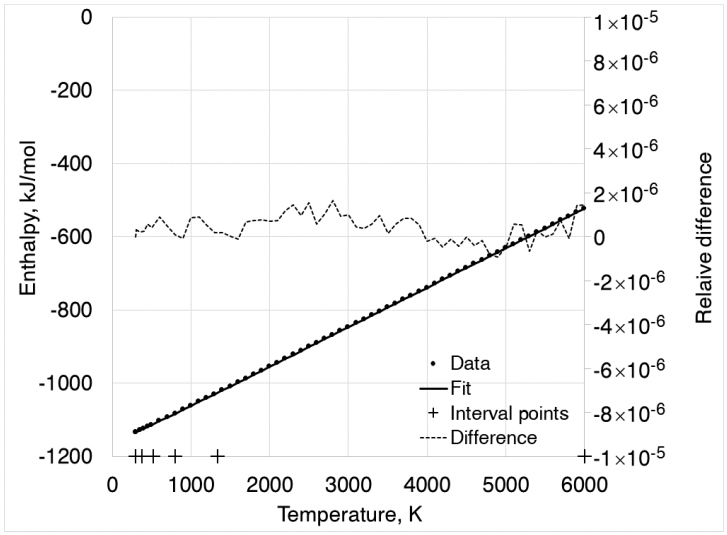


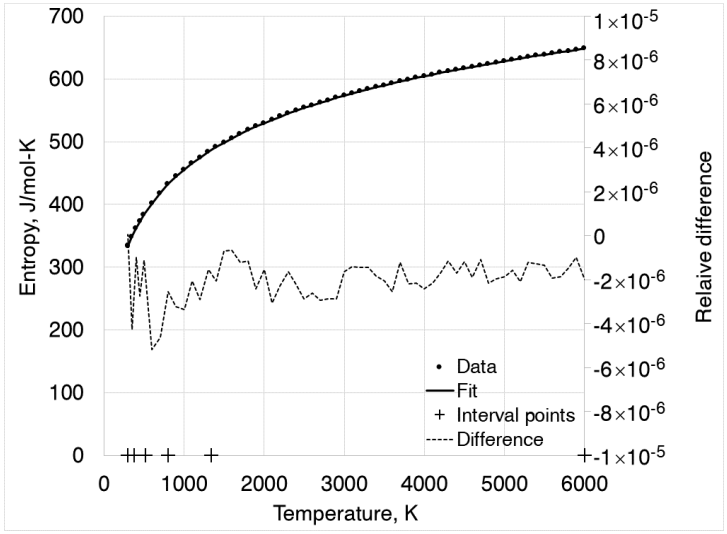


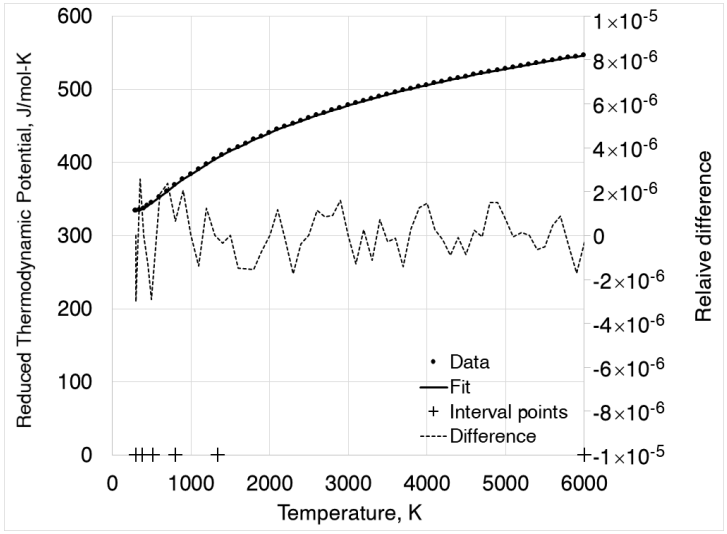
Lead Fluoride (PbF_4) – F4Pb1(g)

MW = 283.2 g/mol – $\Delta_f H^o$(298.15 K) = -1133.446 kJ/mol – S^o(298.15 K) = 333.602 J/mol-K – p^o = 0.1 MPa



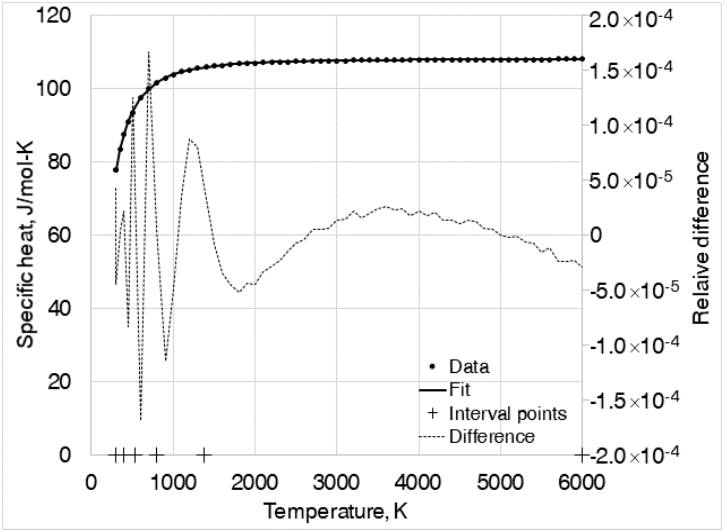


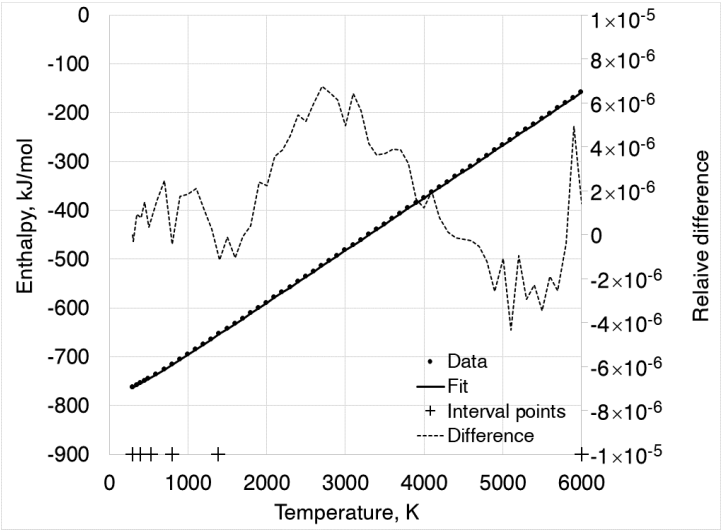


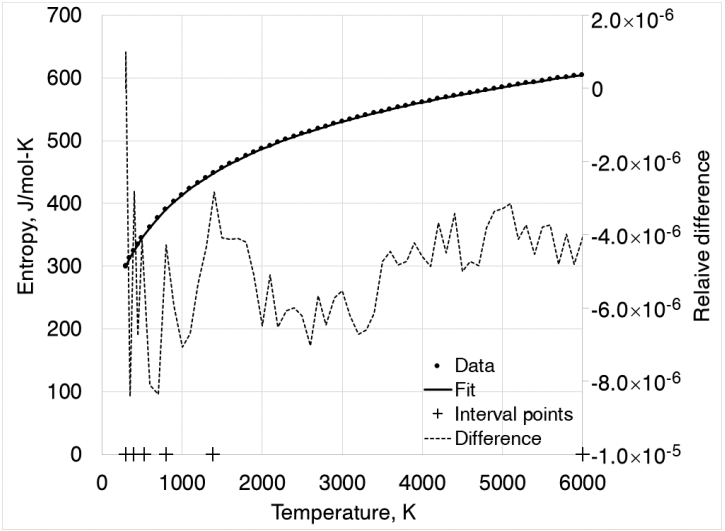


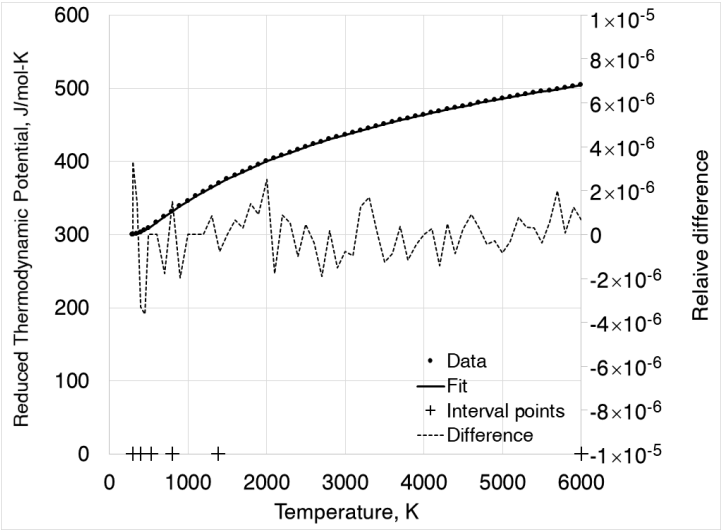
Sulfur Fluoride (SF_4) – F4S1(g)

MW = 108.05 g/mol – $\Delta_f H^o$(298.15 K) = -763.162 kJ/mol – S^o(298.15 K) = 299.644 J/mol-K – p^o = 0.1 MPa



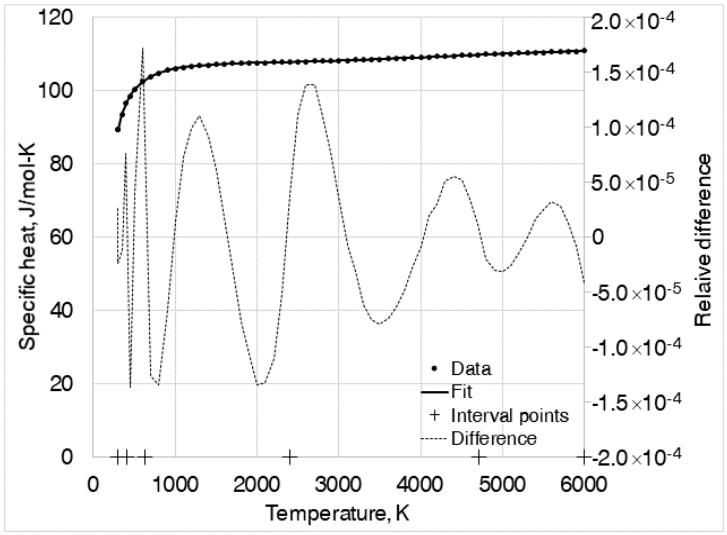


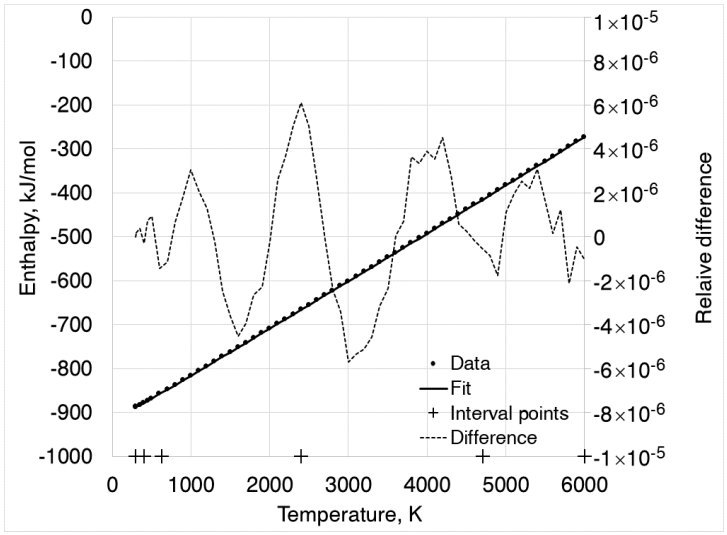


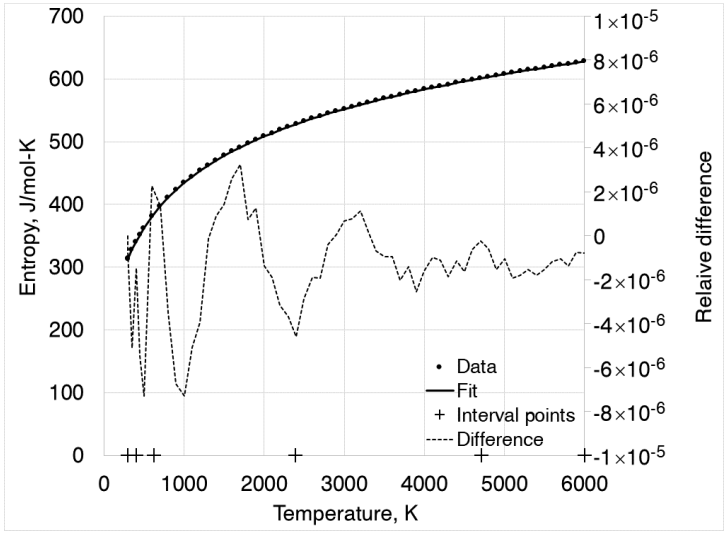


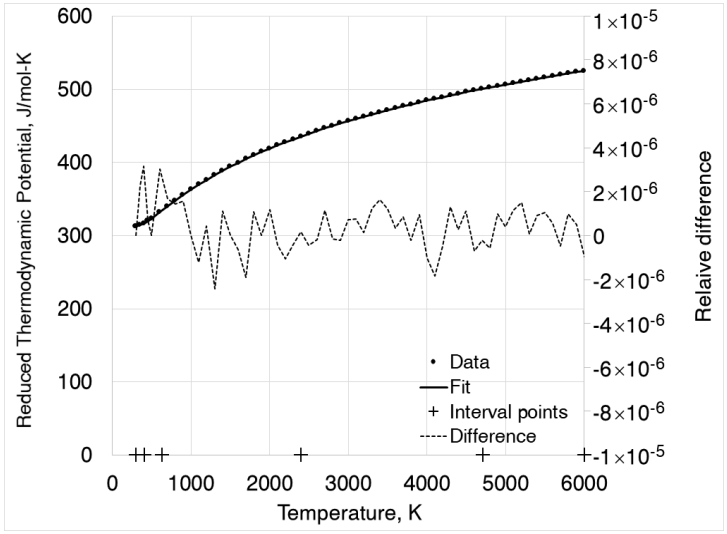
Sulfur Fluoride, Ion (SF_4^-) – F4S1-(g)

MW = 108.05 g/mol – $\Delta_f H^o$(298.15 K) = -887.464 kJ/mol – S^o(298.15 K) = 312.962 J/mol-K – p^o = 0.1 MPa



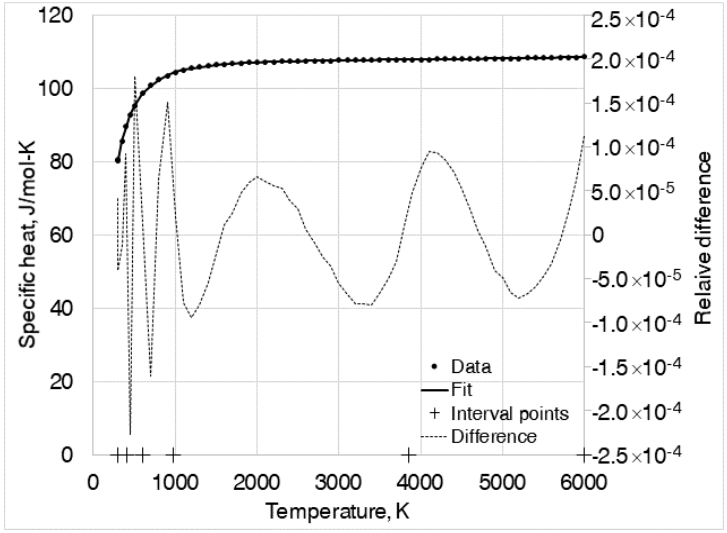


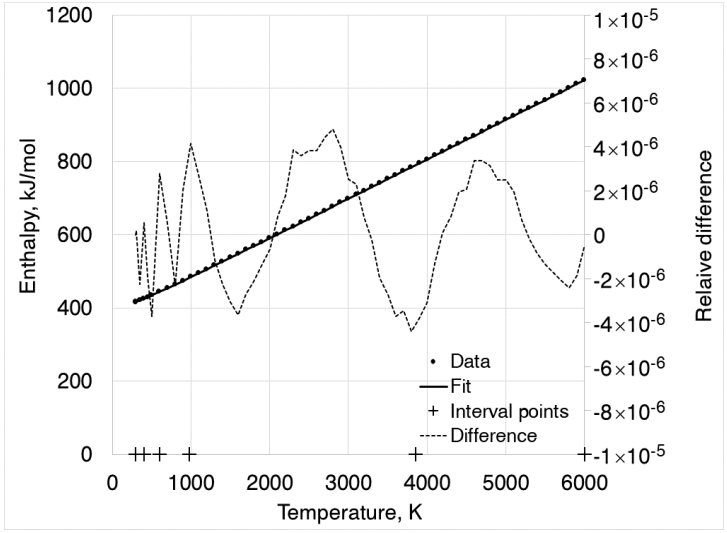


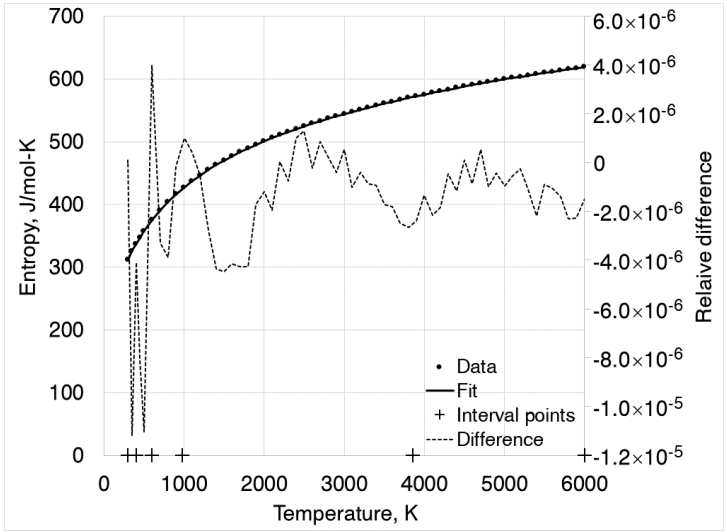


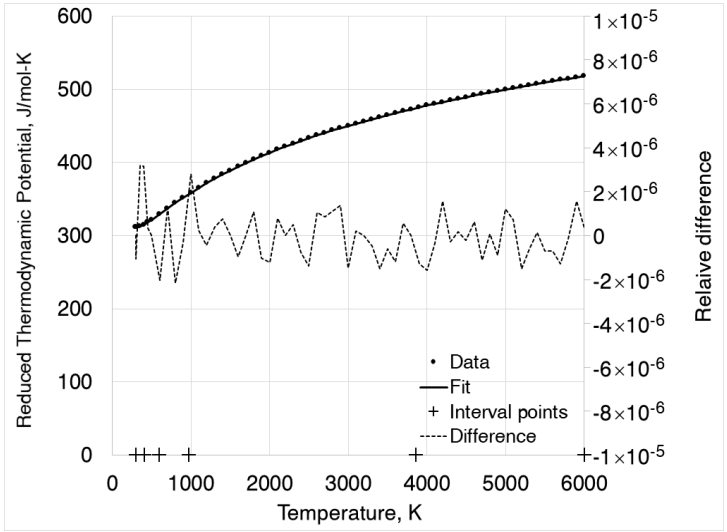
Sulfur Fluoride, Ion (SF_4^+) – F4S1+(g)

MW = 108.05 g/mol – $\Delta_f H^o$(298.15 K) = 416.044 kJ/mol – S^o(298.15 K) = 311.671 J/mol-K – p^o = 0.1 MPa



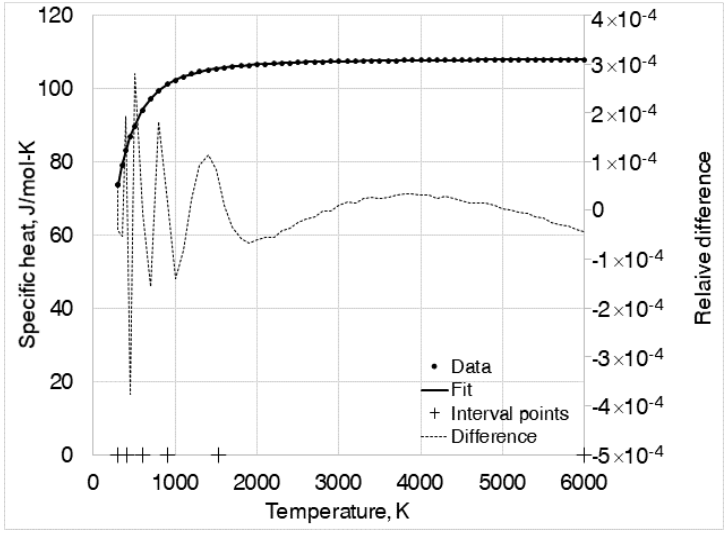


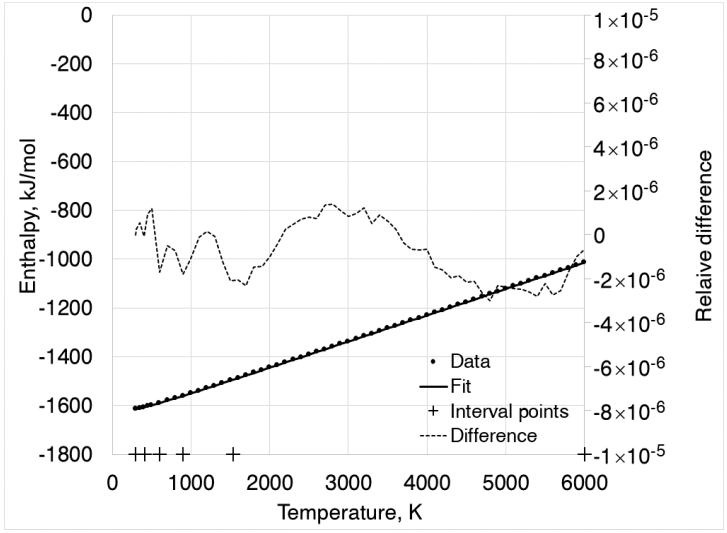


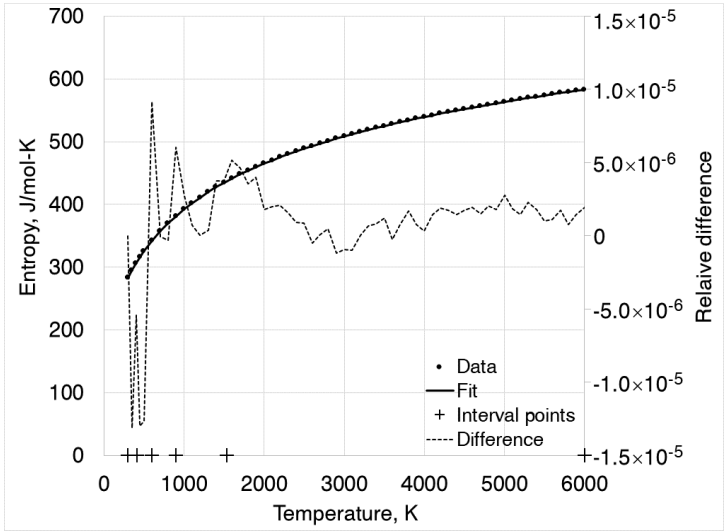


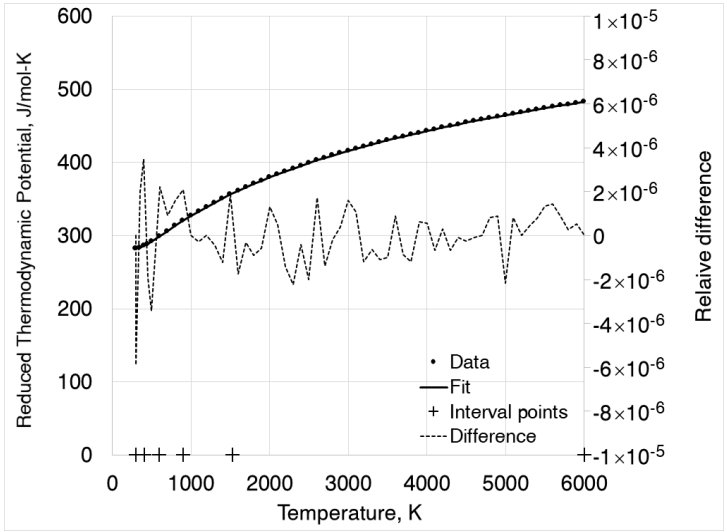
Tetrafluorosilane (SiF_4) – F4Si1(g)

MW = 104.0791 g/mol – $\Delta_f H^o$(298.15 K) = -1614.94 kJ/mol – S^o(298.15 K) = 282.76 J/mol-K – p^o = 0.1 MPa



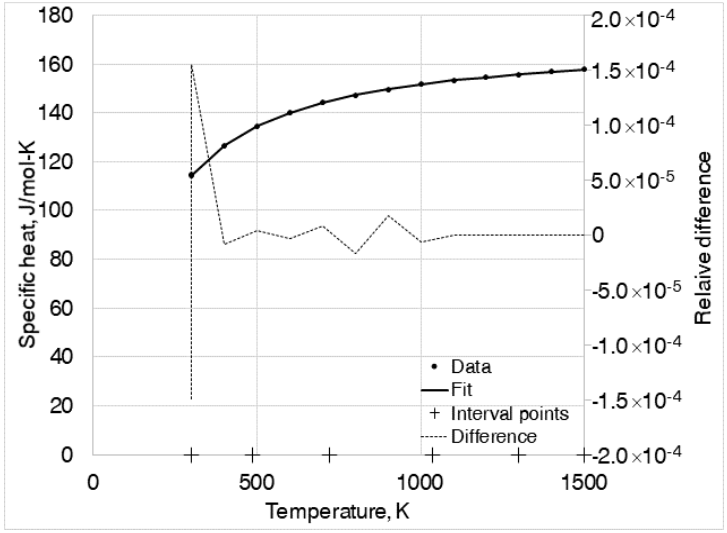


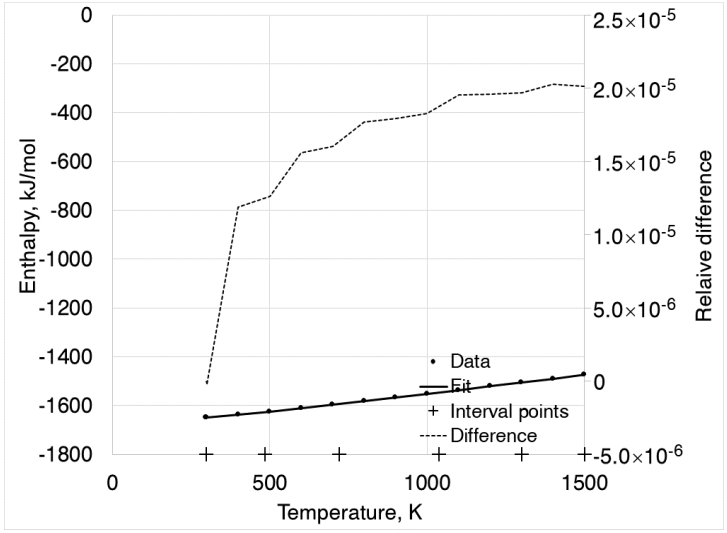


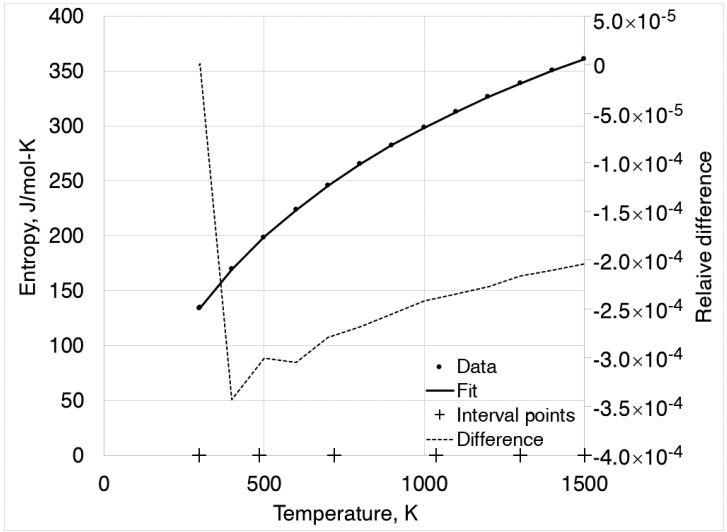


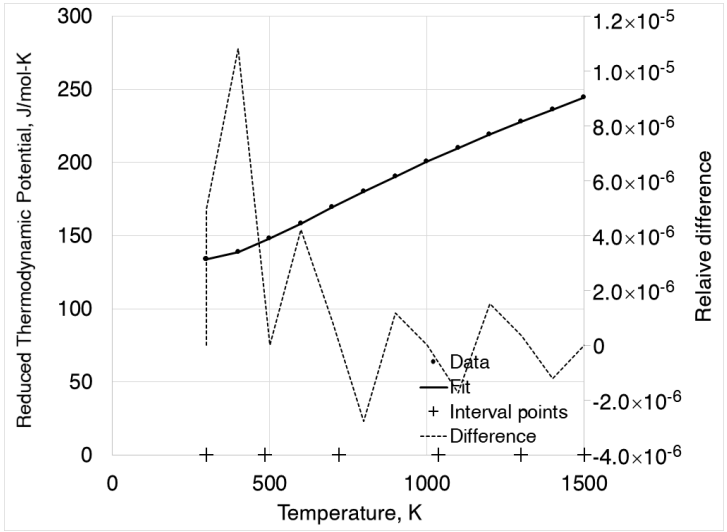
Titanium Fluoride (TiF_4) – F4Ti1(cr)

MW = 123.87 g/mol – $\Delta_f H^o$(298.15 K) = -1649.333 kJ/mol – S^o(298.15 K) = 133.955 J/mol-K – p^o = 0.1 MPa



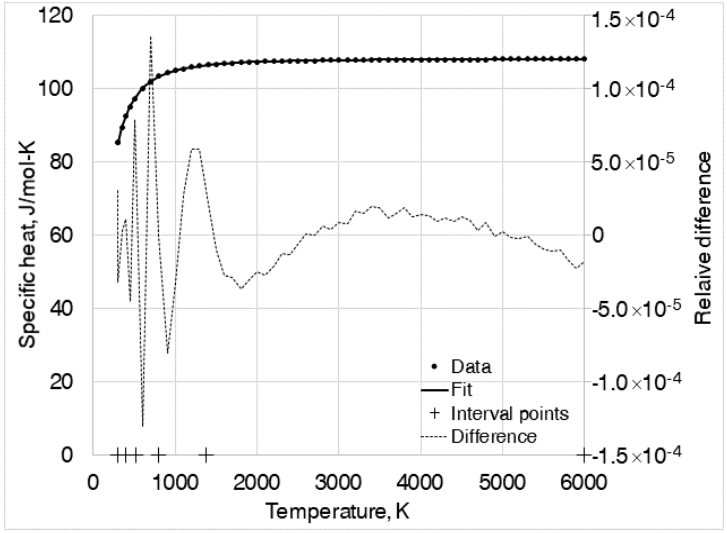


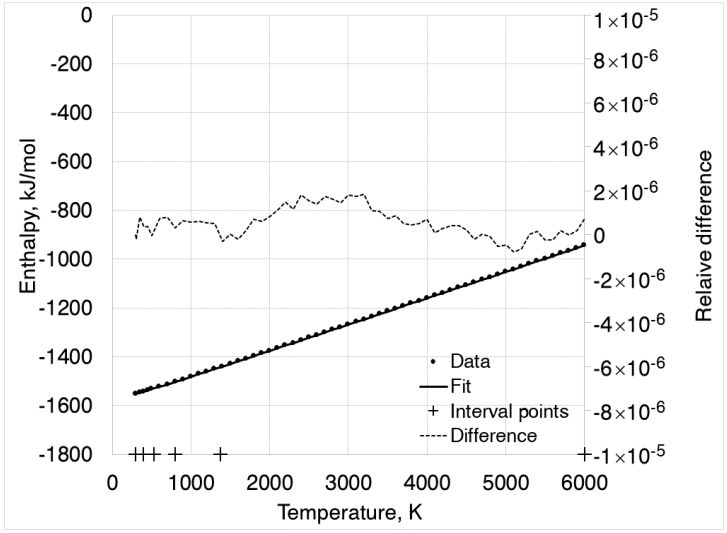


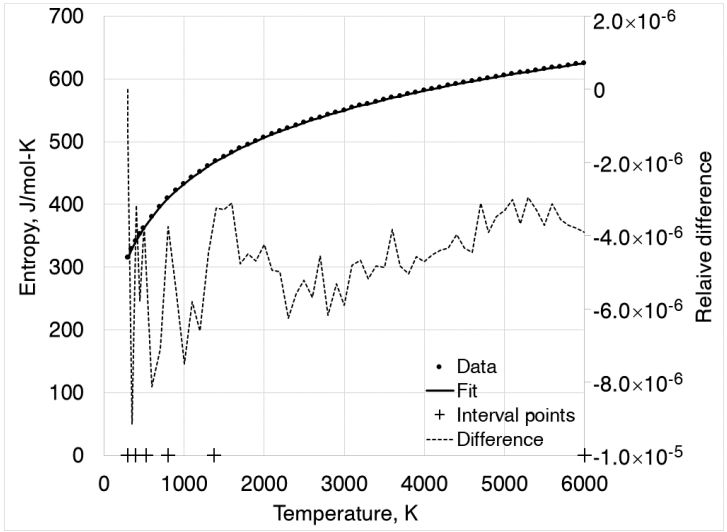


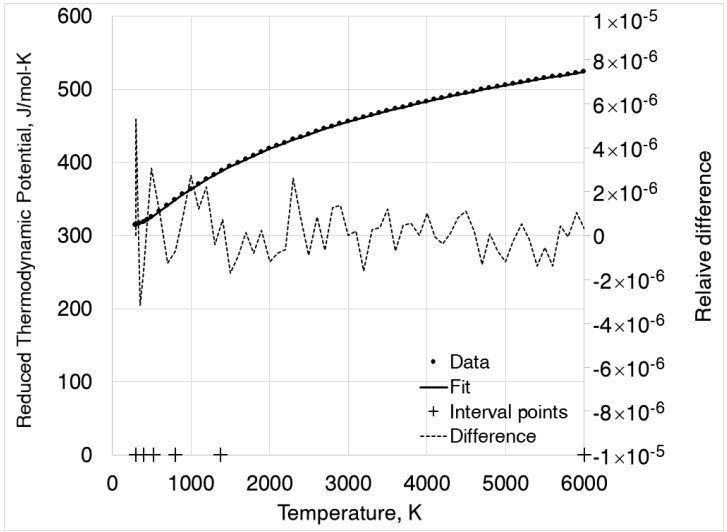
Titanium Fluoride (TiF_4) – F4Ti1(g)

MW = 123.87 g/mol – $\Delta_f H^o$(298.15 K) = -1551.427 kJ/mol – S^o(298.15 K) = 314.928 J/mol-K – p^o = 0.1 MPa



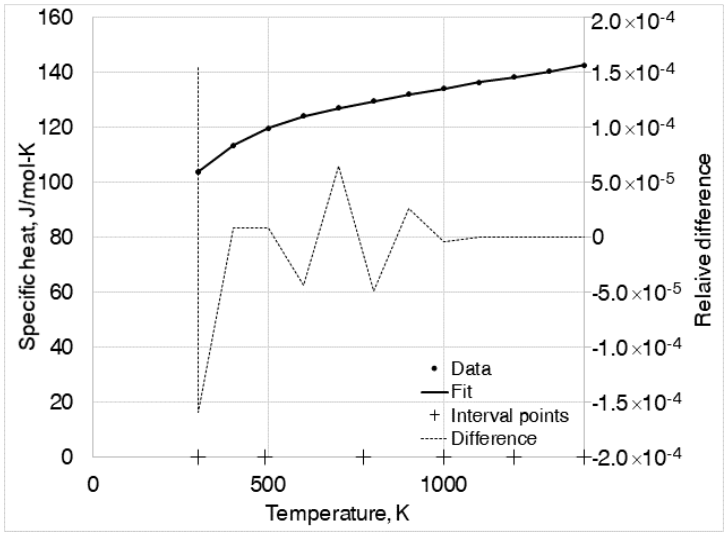


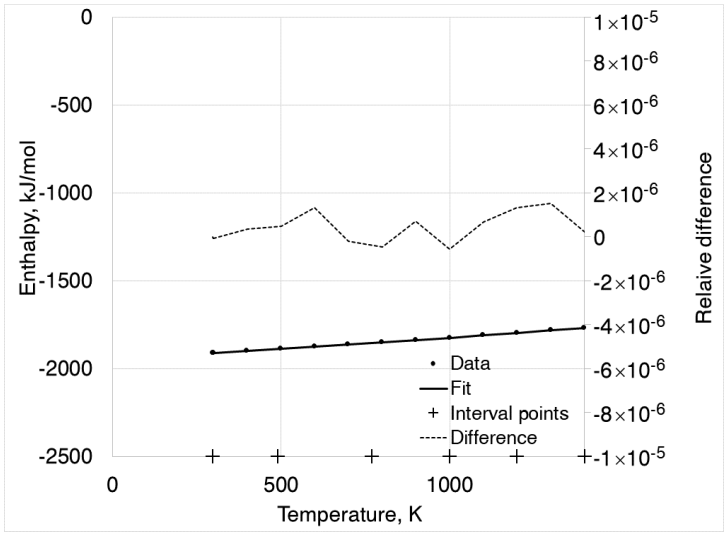


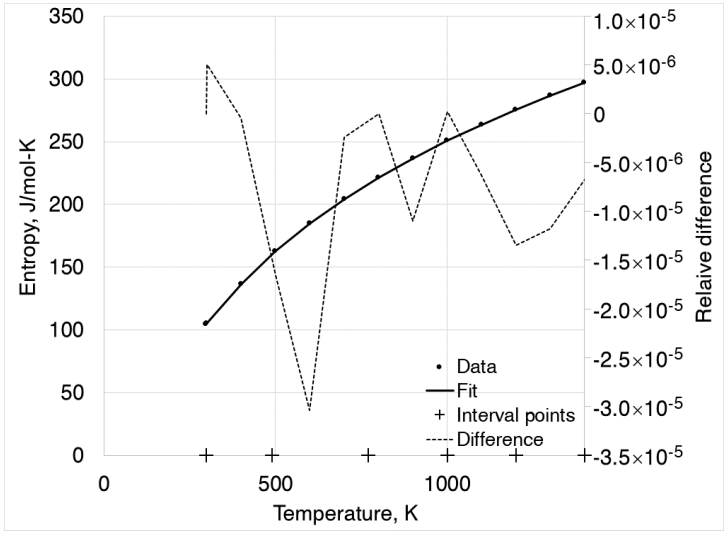


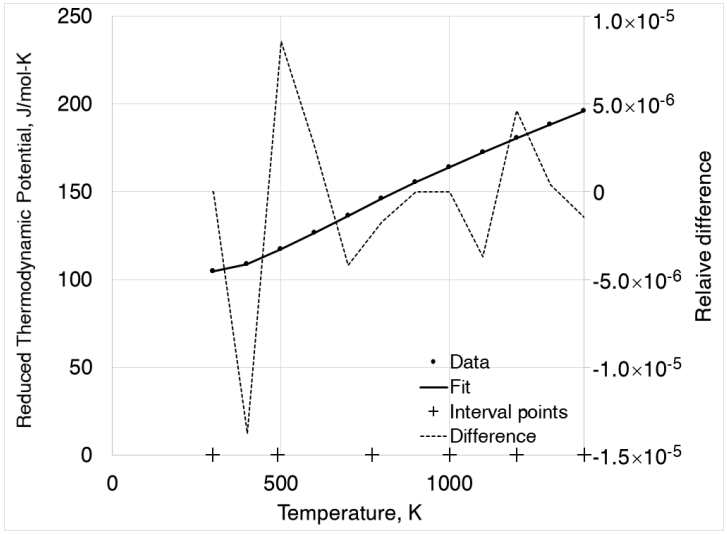
Zirconium Fluoride (ZrF_4) – F4Zr1(cr)

MW = 167.21 g/mol – $\Delta_f H^o$(298.15 K) = -1911.251 kJ/mol – S^o(298.15 K) = 104.7 J/mol-K – p^o = 0.1 MPa



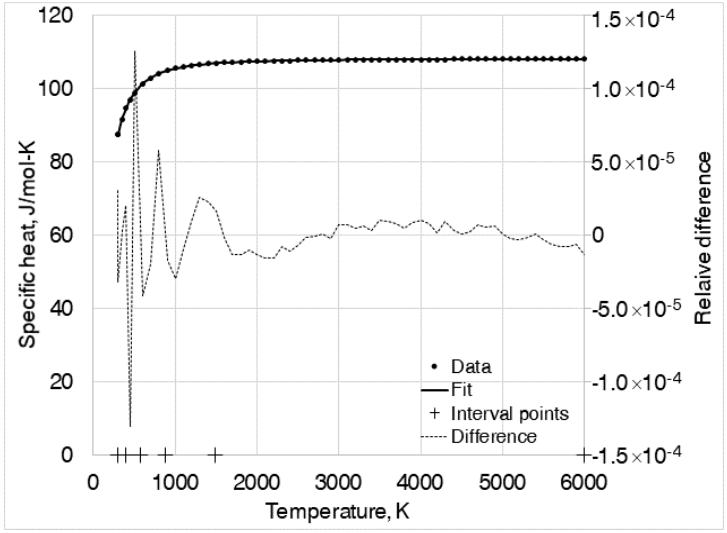


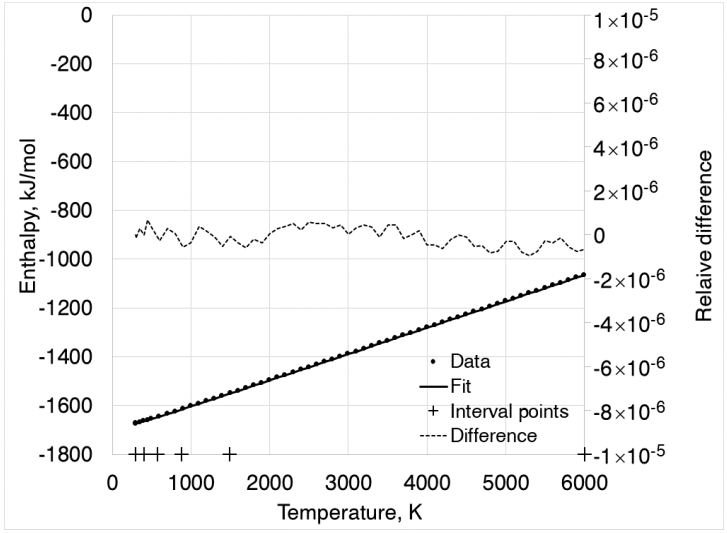


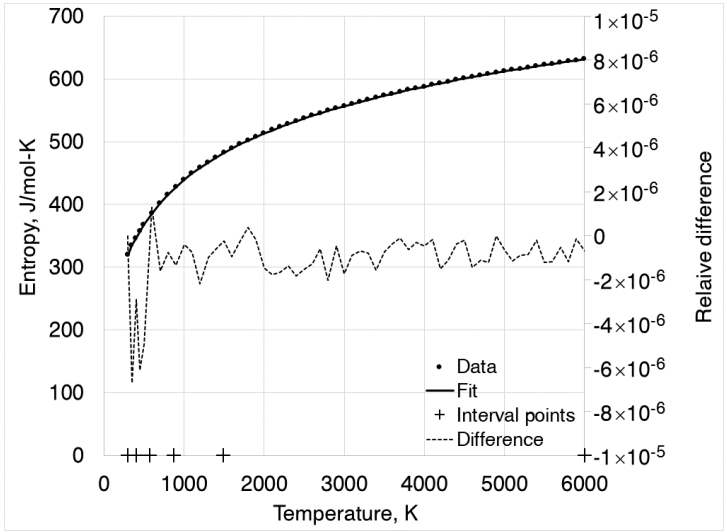


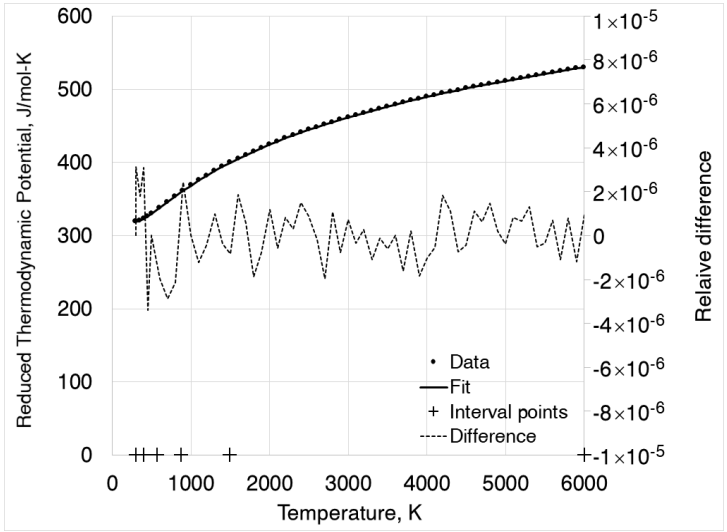
Zirconium Fluoride (ZrF_4) – F4Zr1(g)

MW = 167.21 g/mol – $\Delta_f H^o$(298.15 K) = -1673.6 kJ/mol – S^o(298.15 K) = 319.516 J/mol-K – p^o = 0.1 MPa



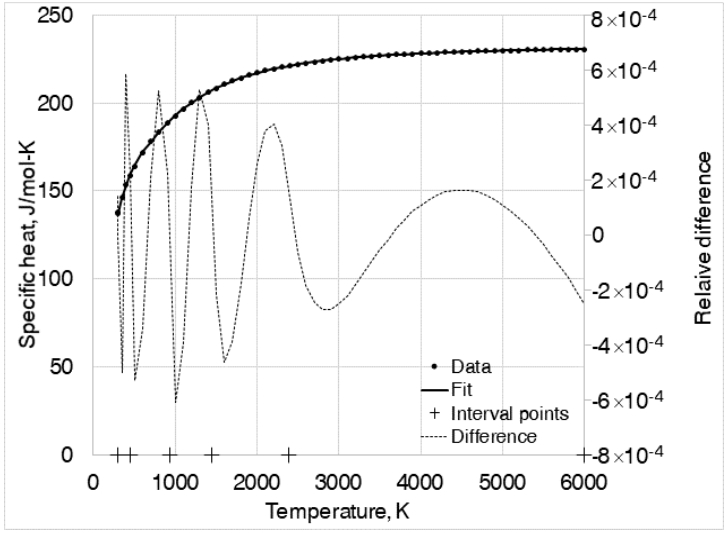


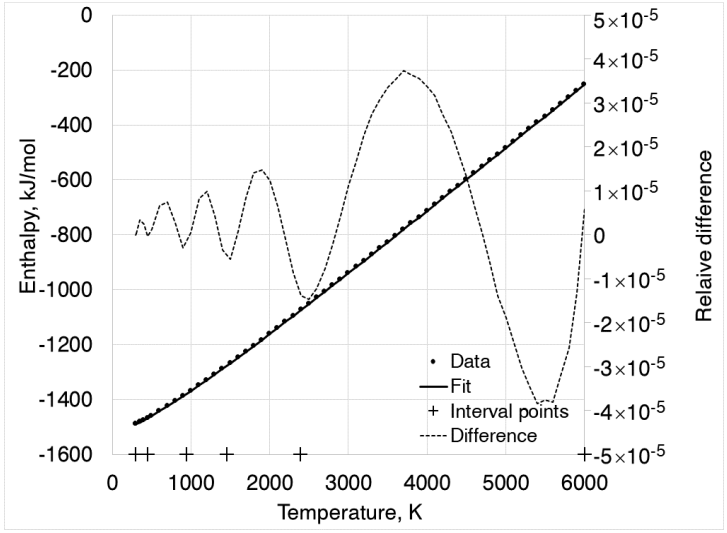


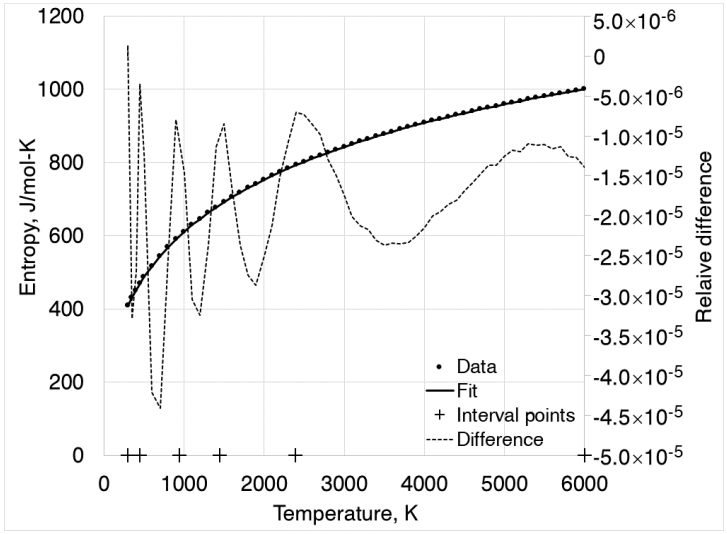


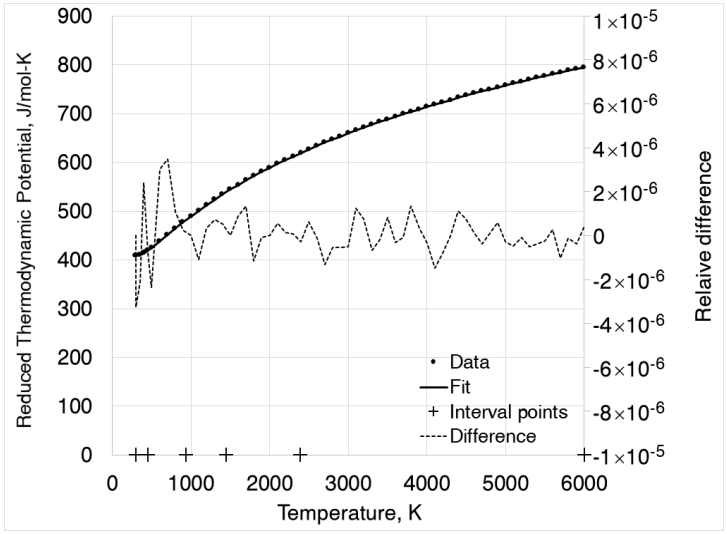
Hydrogen Fluoride (H_5F_5) – F5H5(g)

MW = 100.0317 g/mol – $\Delta_f H^o$(298.15 K) = -1489.713 kJ/mol – S^o(298.15 K) = 408.442 J/mol-K – p^o = 0.1 MPa



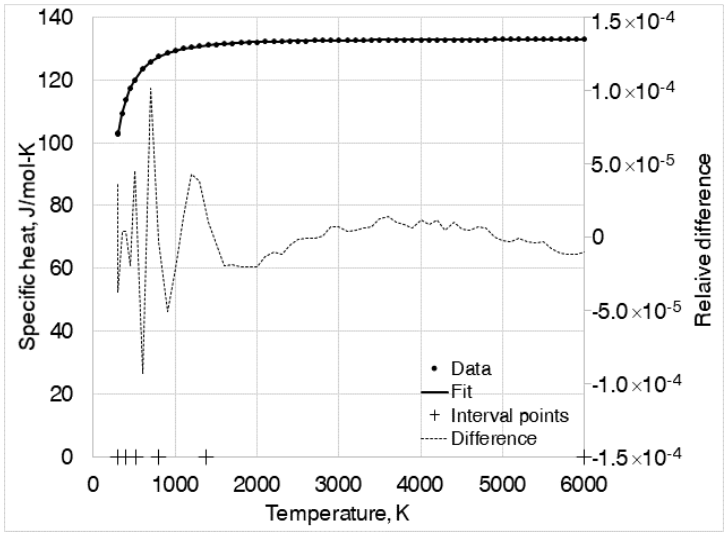


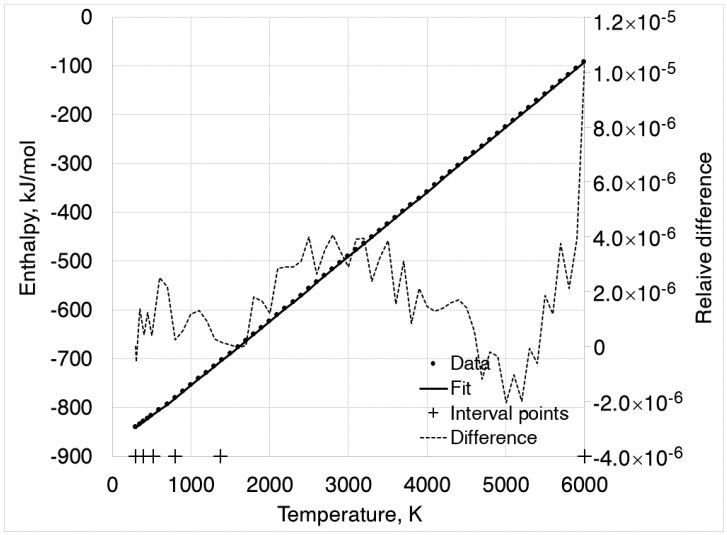


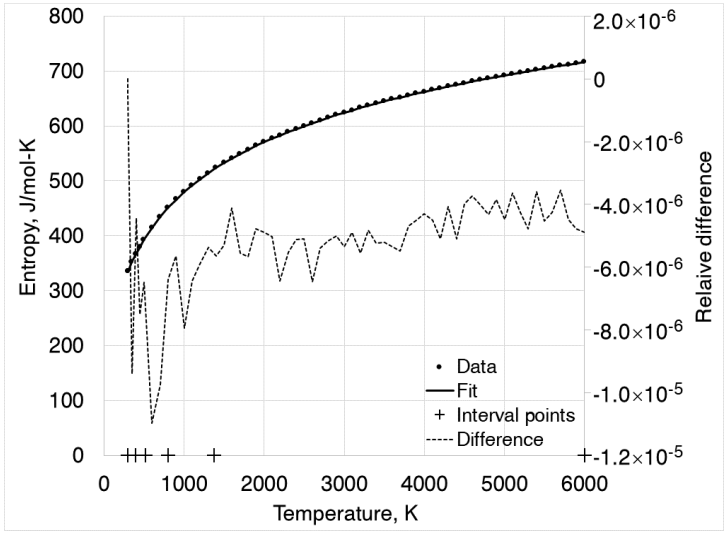


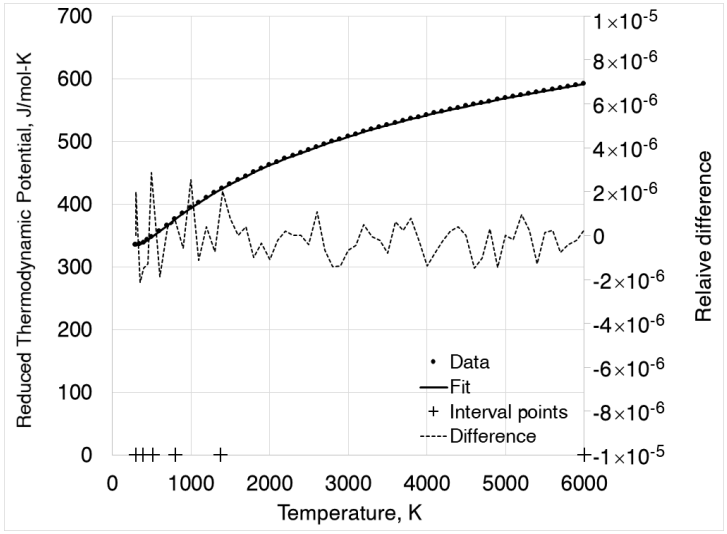
Iodine Fluoride (IF_5) – F5I1(g)

MW = 221.8965 g/mol – $\Delta_f H^o$(298.15 K) = -840.315 kJ/mol – S^o(298.15 K) = 334.663 J/mol-K – p^o = 0.1 MPa



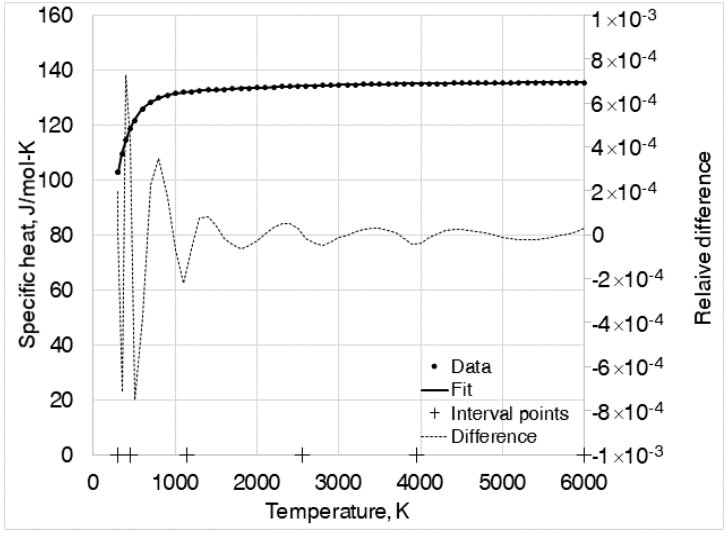


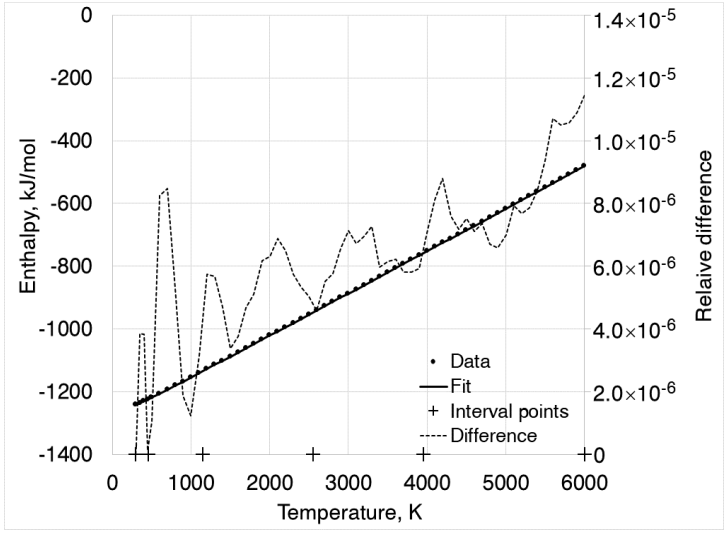


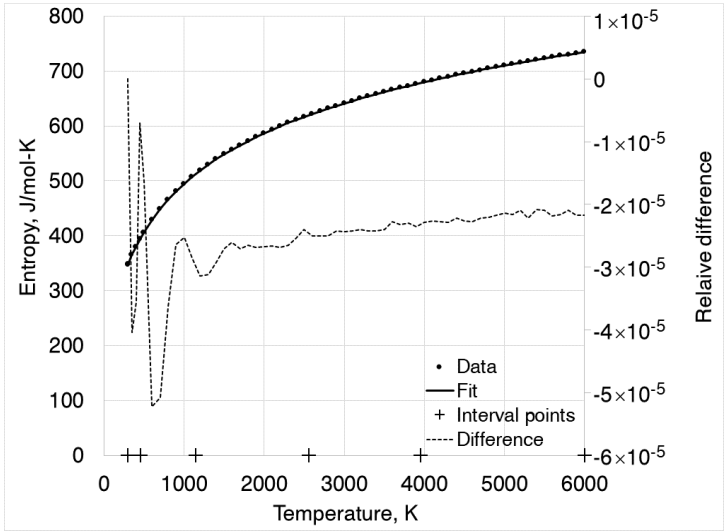


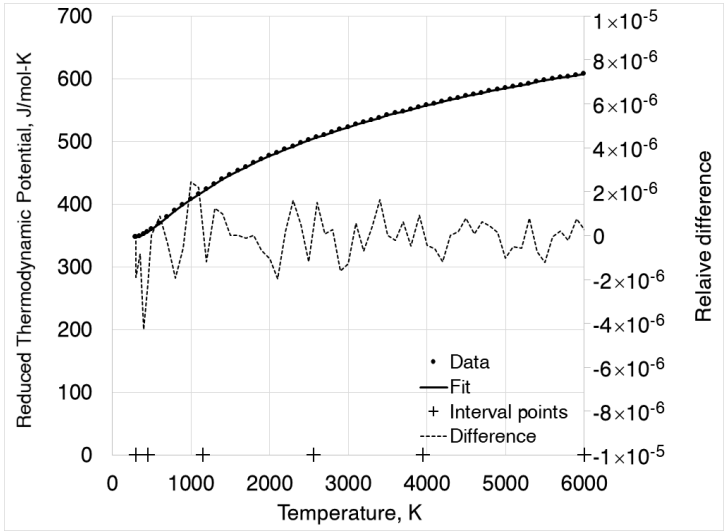
Molybdenum Fluoride (MoF_5) – F5Mo1(g)

MW = 190.93 g/mol – $\Delta_f H^o$(298.15 K) = -1241.393 kJ/mol – S^o(298.15 K) = 347.655 J/mol-K – p^o = 0.1 MPa



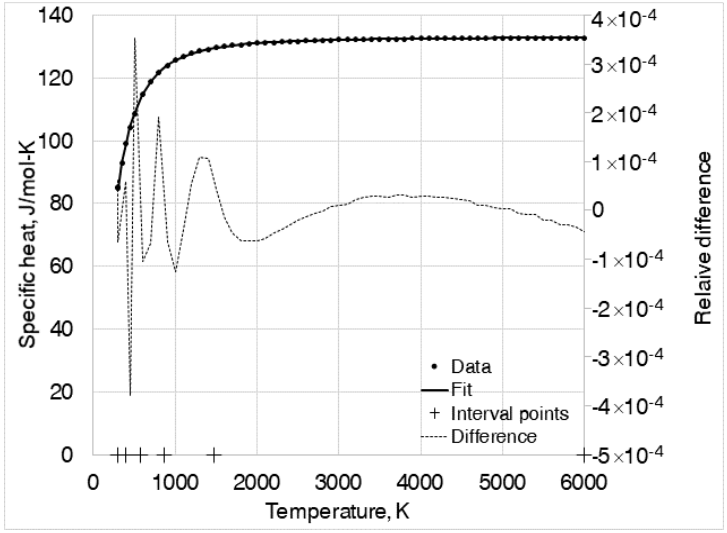


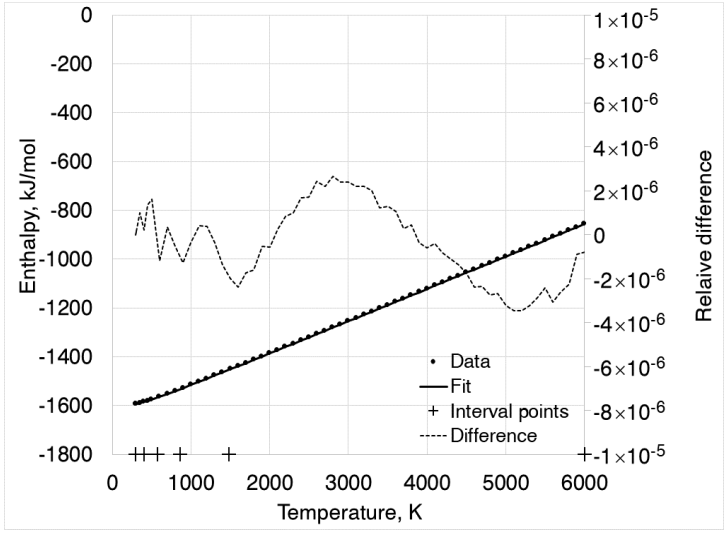


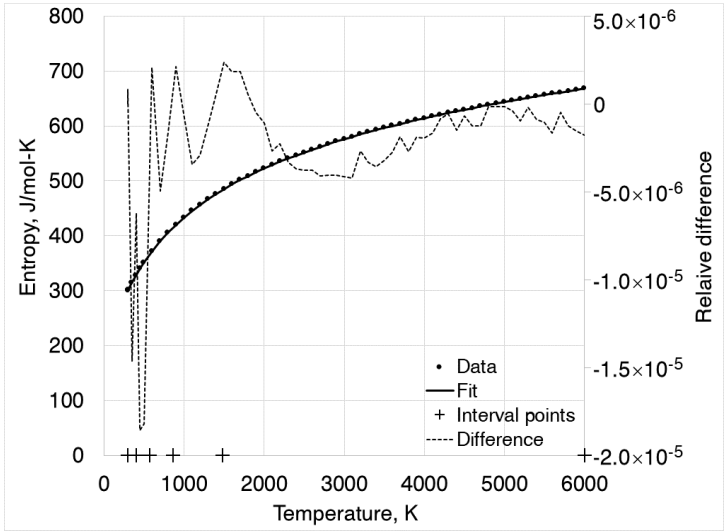


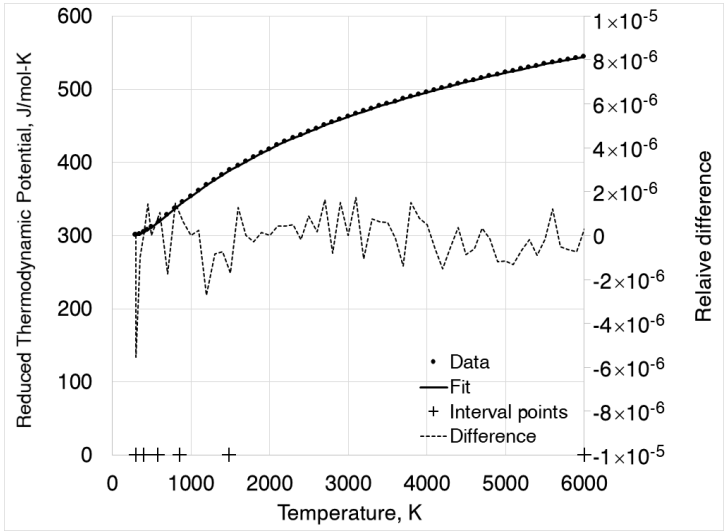
Phosphorus Fluoride (PF_5) – F5P1(g)

MW = 125.9658 g/mol – $\Delta_f H^o$(298.15 K) = -1594.409 kJ/mol – S^o(298.15 K) = 300.81 J/mol-K – p^o = 0.1 MPa



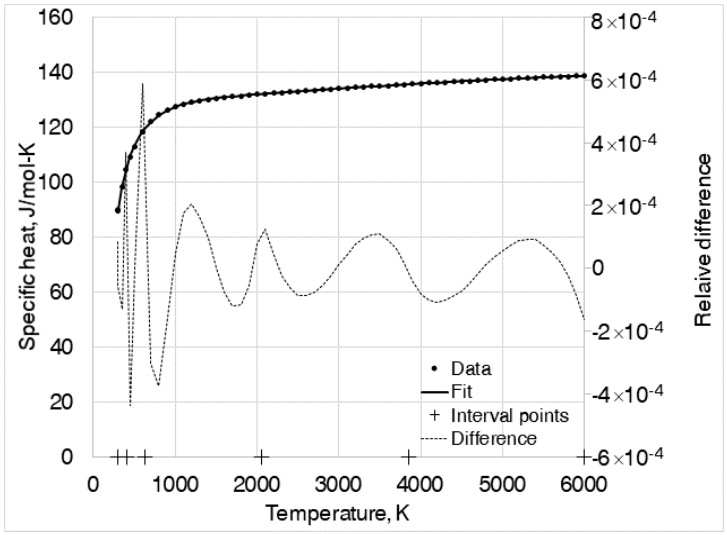


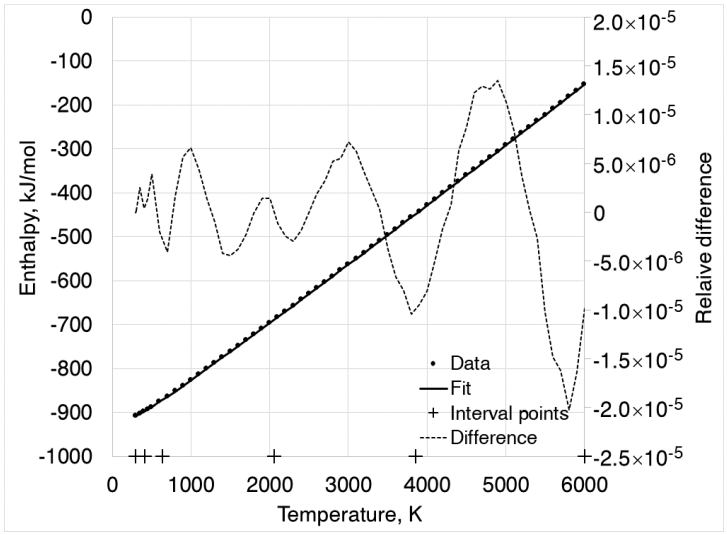


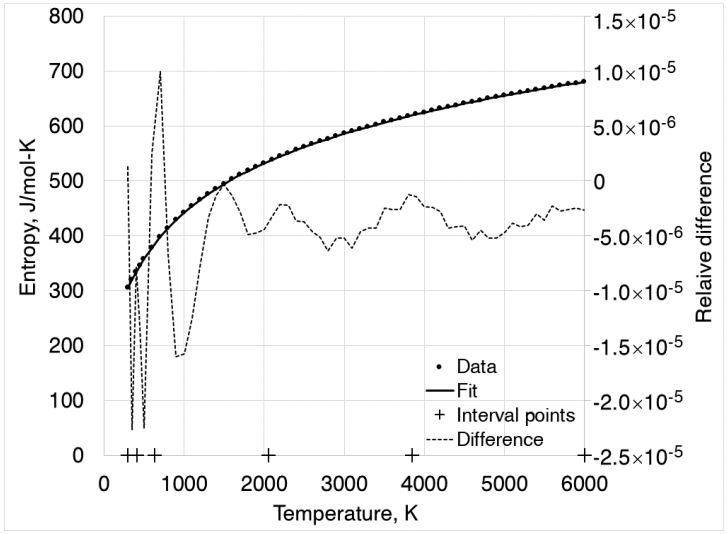


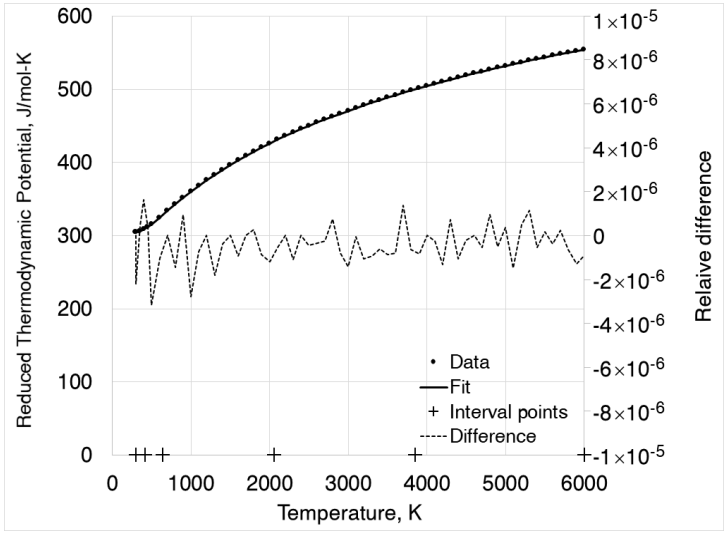
Sulfur Fluoride (SF_5) – F5S1(g)

MW = 127.05 g/mol – $\Delta_f H^o$(298.15 K) = -908.447 kJ/mol – S^o(298.15 K) = 304.752 J/mol-K – p^o = 0.1 MPa



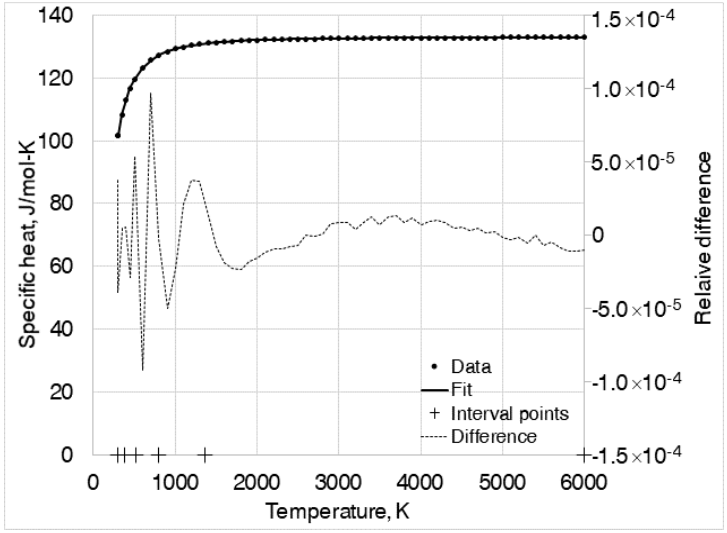


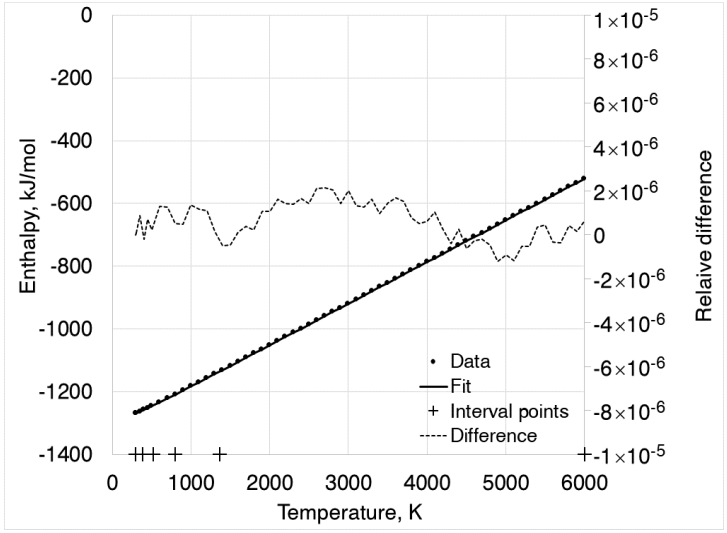


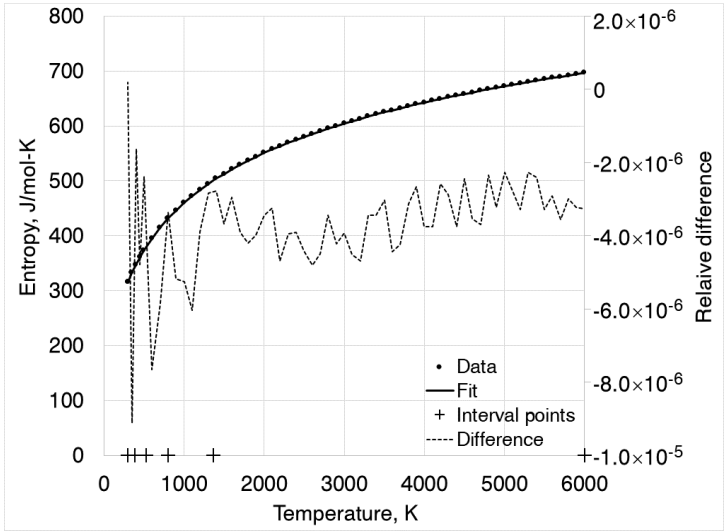


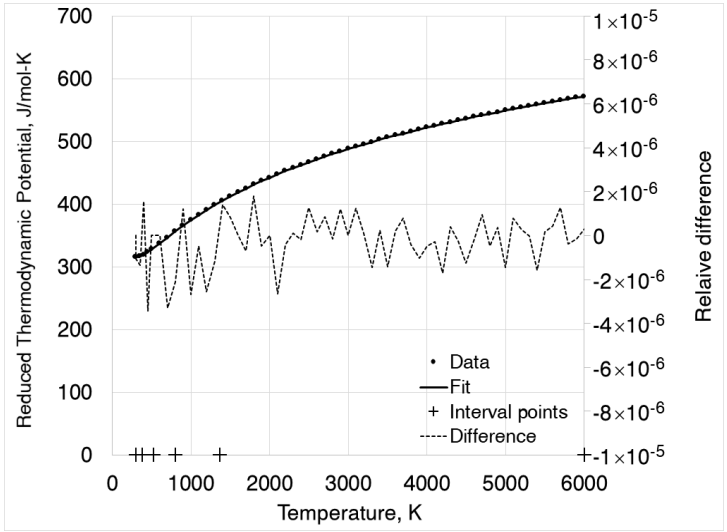
Sulfur Fluoride, Ion (SF_5^-) – F5S1-(g)

MW = 127.05 g/mol – $\Delta_f H^o$(298.15 K) = -1268.706 kJ/mol – S^o(298.15 K) = 315.571 J/mol-K – p^o = 0.1 MPa



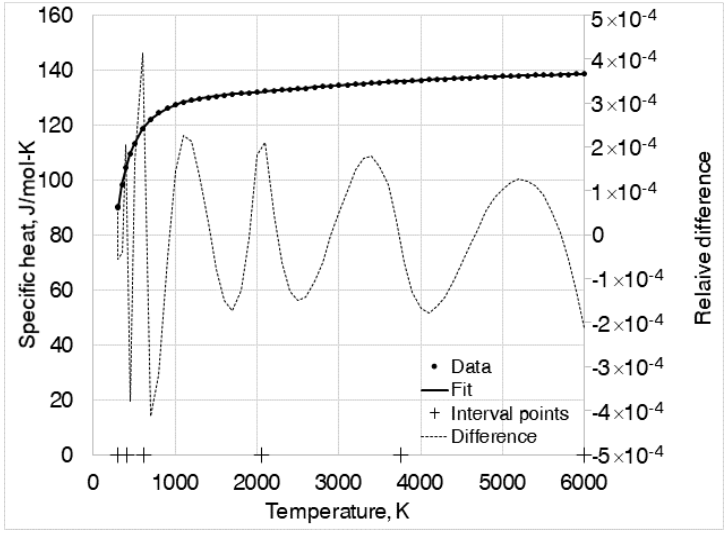


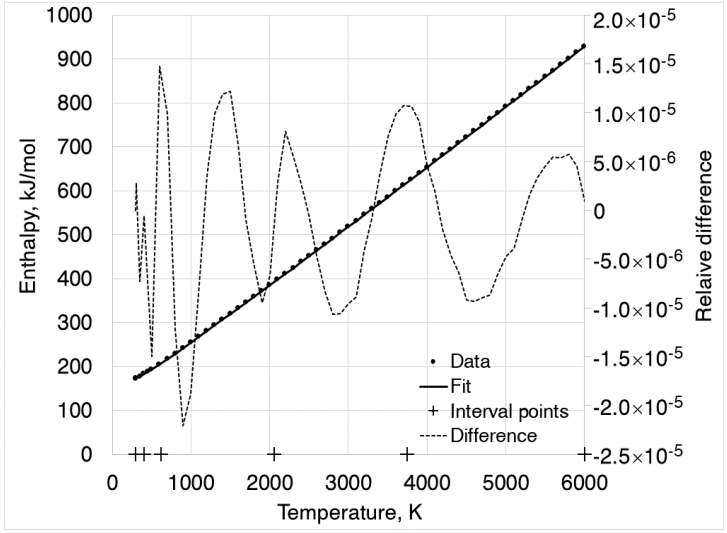


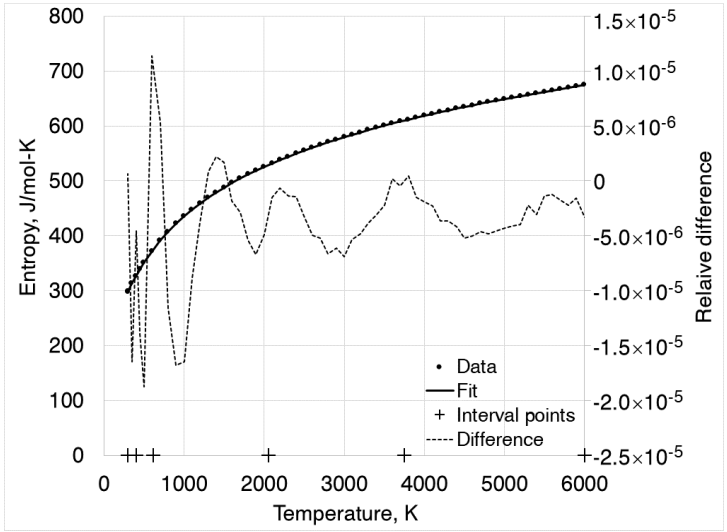


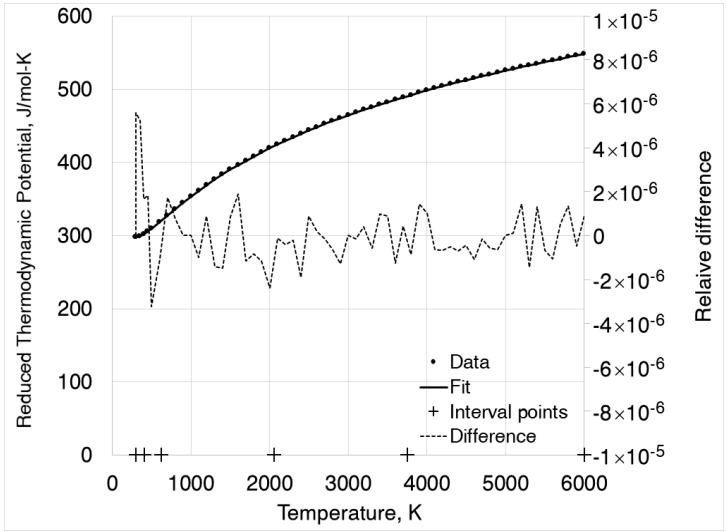
Sulfur Fluoride, Ion (SF_5^+) – F5S1+(g)

MW = 127.05 g/mol – $\Delta_f H^o$(298.15 K) = 172.644 kJ/mol – S^o(298.15 K) = 298.151 J/mol-K – p^o = 0.1 MPa



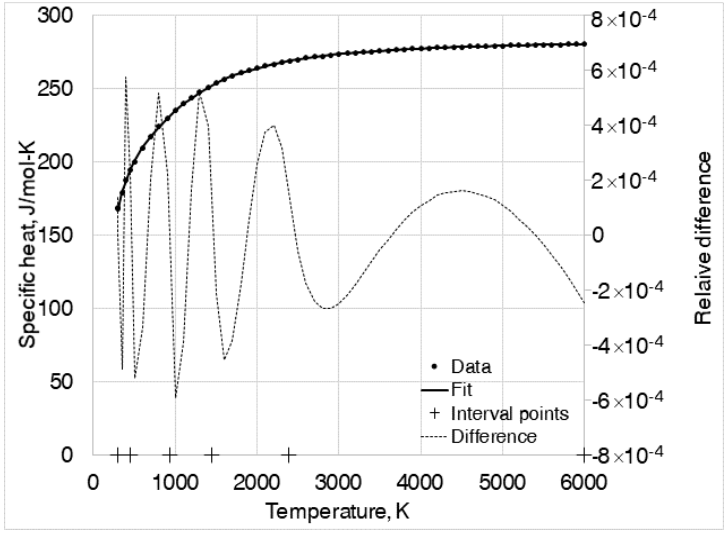


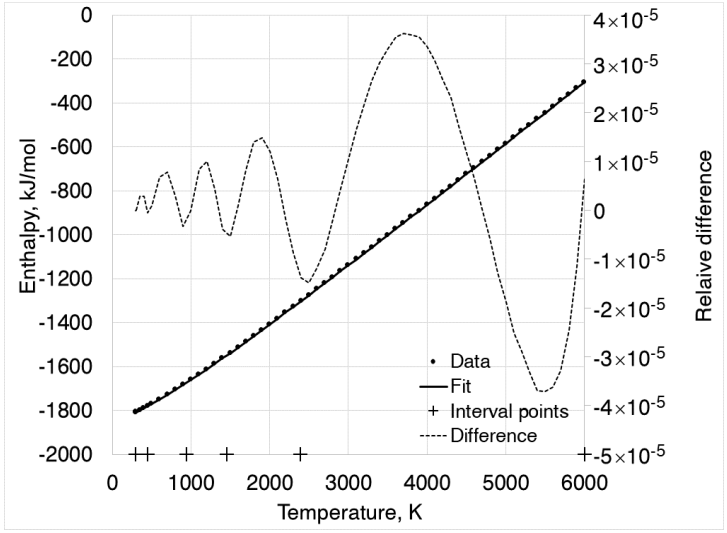


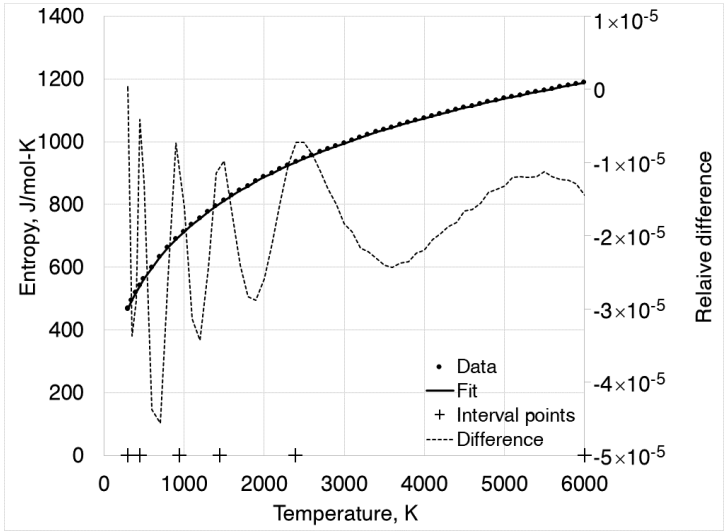


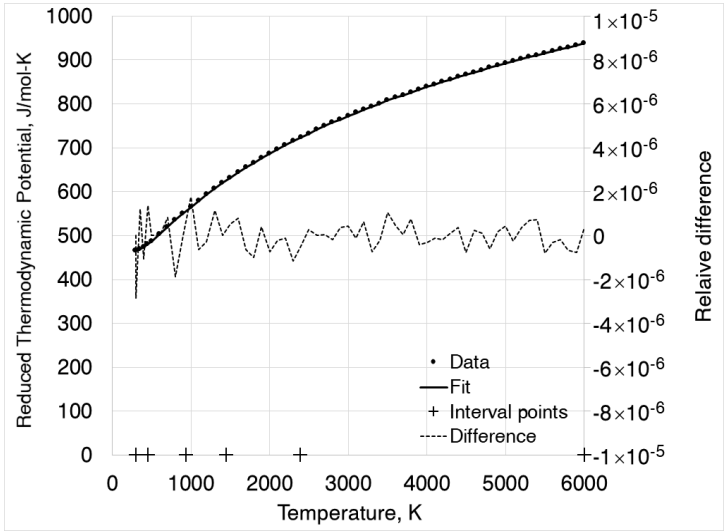
Hydrogen Fluoride (H_6F_6) – F6H6(g)

MW = 120.0381 g/mol – $\Delta_f H^o$(298.15 K) = -1807.655 kJ/mol – S^o(298.15 K) = 466.42 J/mol-K – p^o = 0.1 MPa



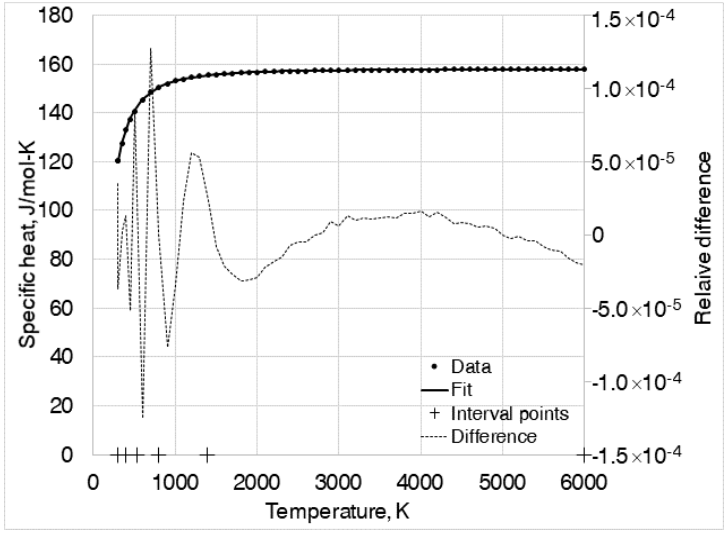


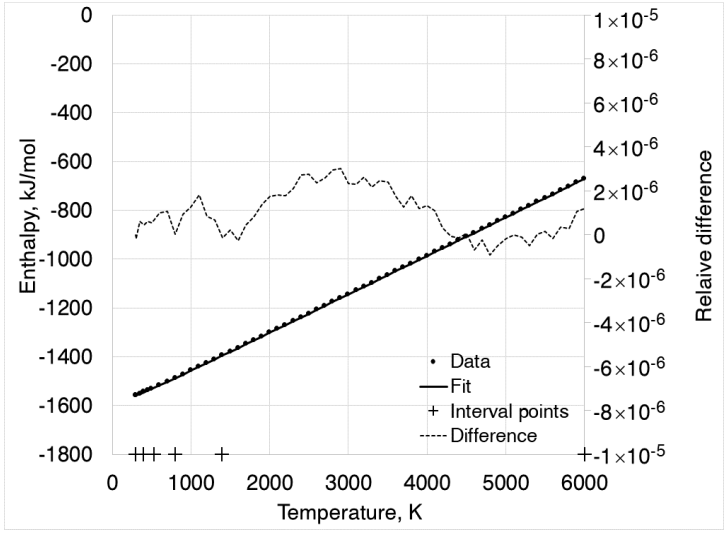


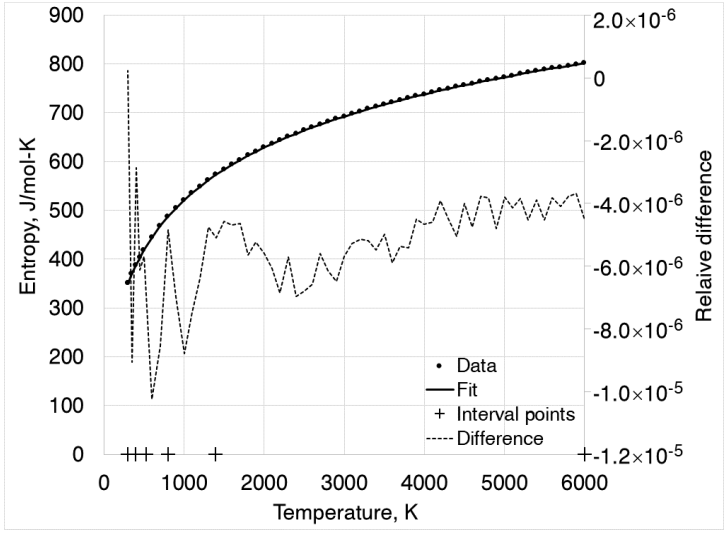


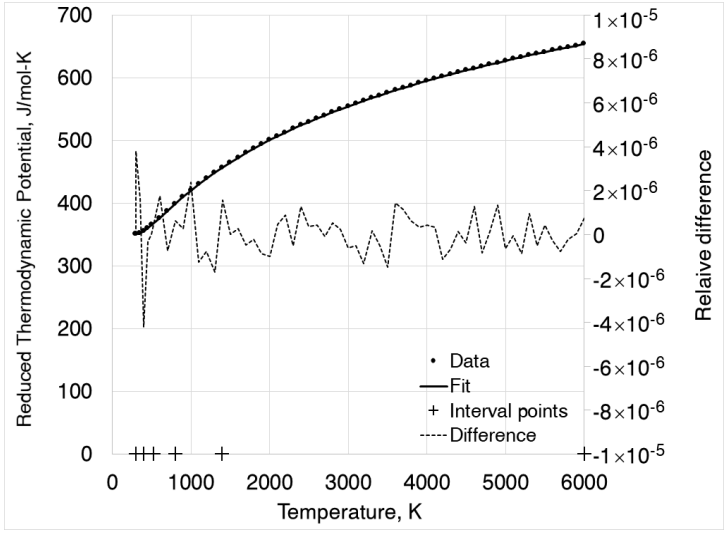
Molybdenum Fluoride (MoF_6) – F6Mo1(g)

MW = 209.93 g/mol – $\Delta_f H^o$(298.15 K) = -1557.661 kJ/mol – S^o(298.15 K) = 350.71 J/mol-K – p^o = 0.1 MPa



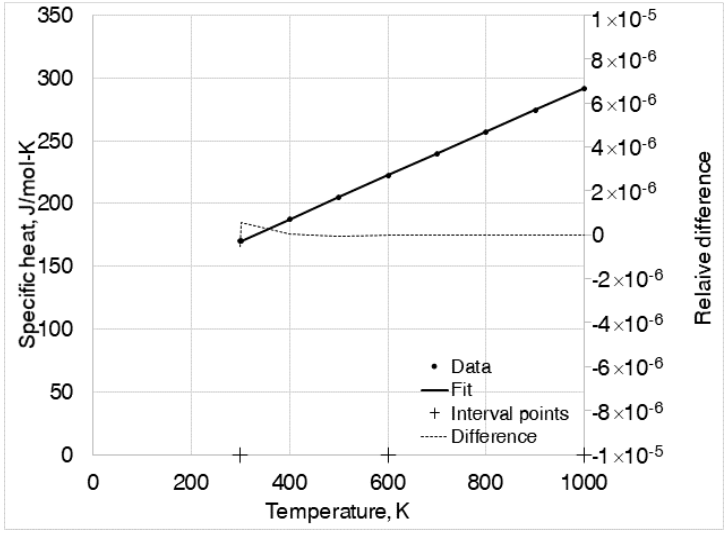


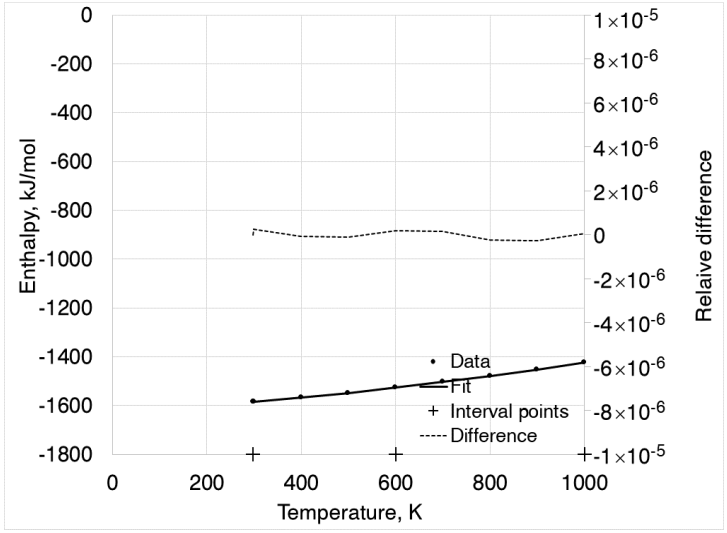


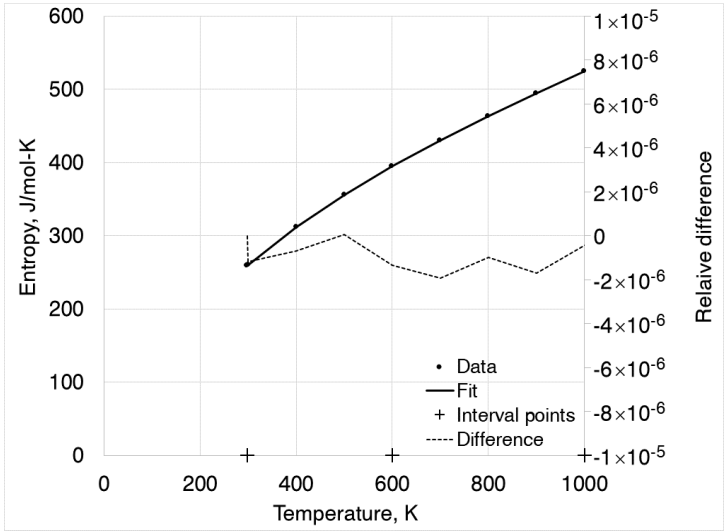


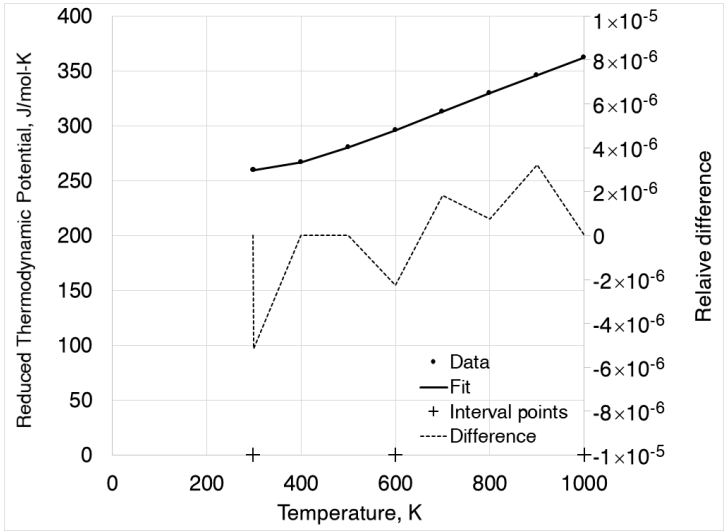
Molybdenum Fluoride (MoF_6) – F6Mo1(l)

MW = 209.93 g/mol – $\Delta_f H^o$(298.15 K) = -1585.656 kJ/mol – S^o(298.15 K) = 259.693 J/mol-K – p^o = 0.1 MPa



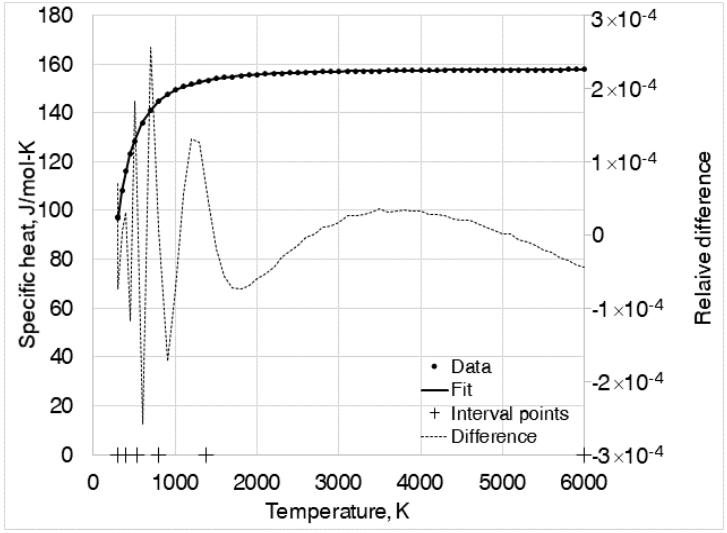


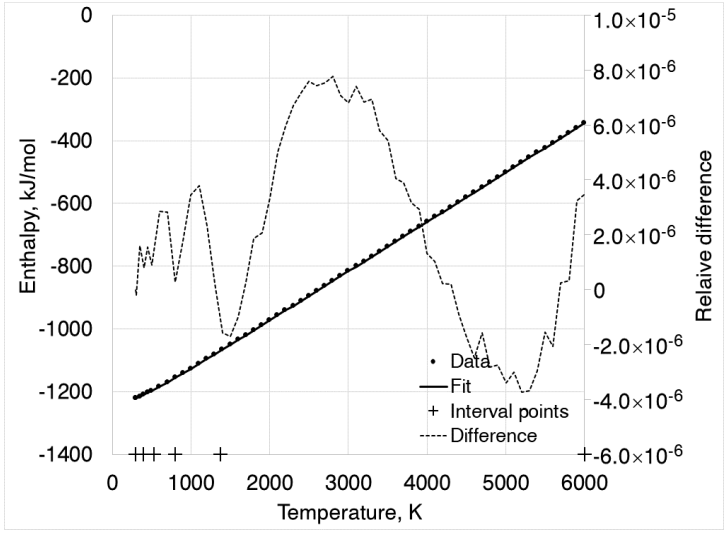


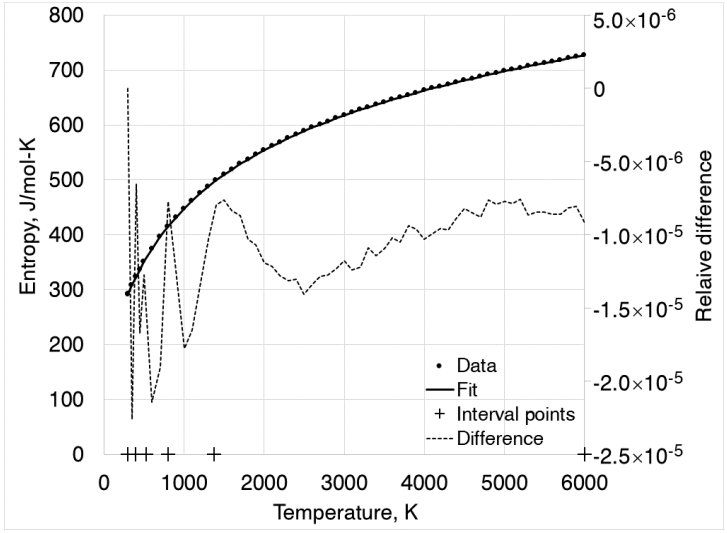


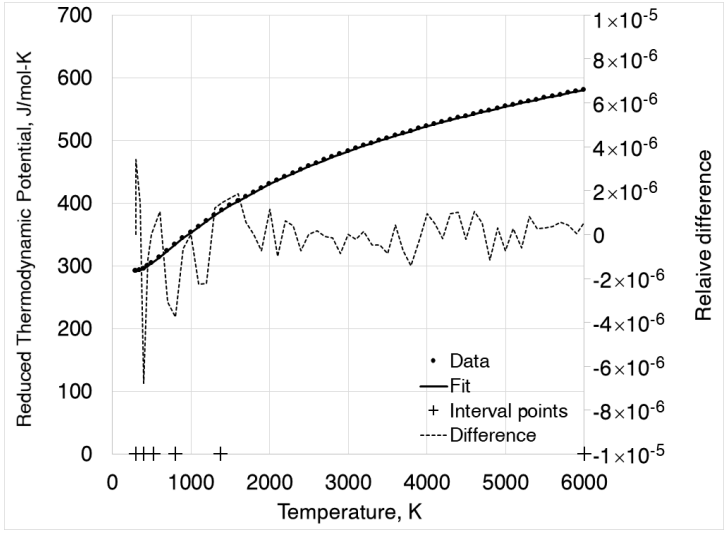
Sulfur Fluoride (SF_6) – F6S1(g)

MW = 146.05 g/mol – $\Delta_f H^o$(298.15 K) = -1220.473 kJ/mol – S^o(298.15 K) = 291.535 J/mol-K – p^o = 0.1 MPa



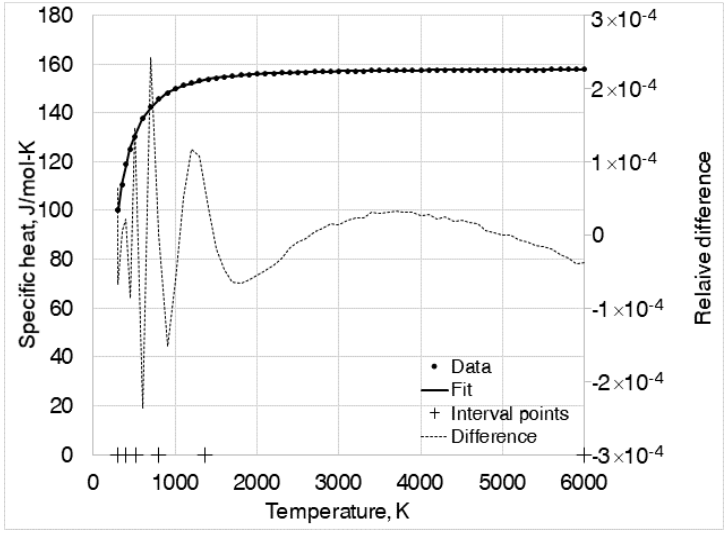


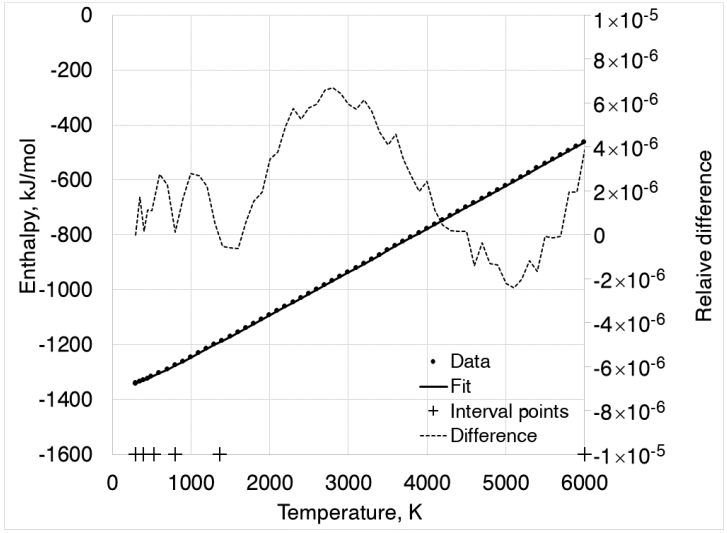


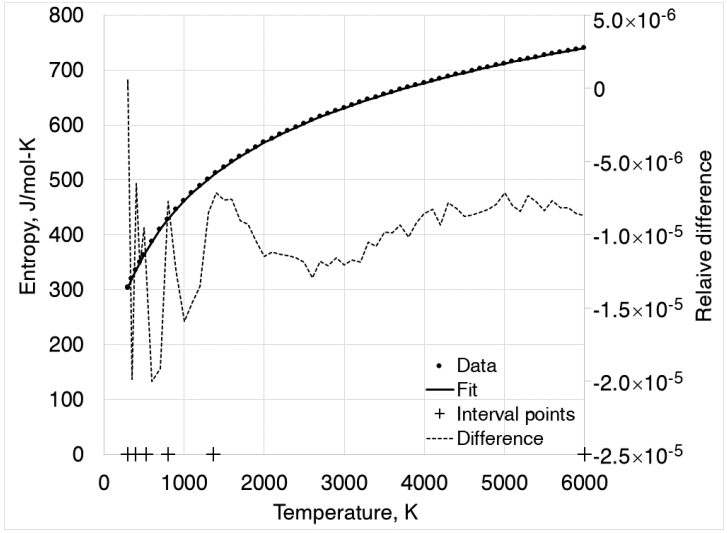


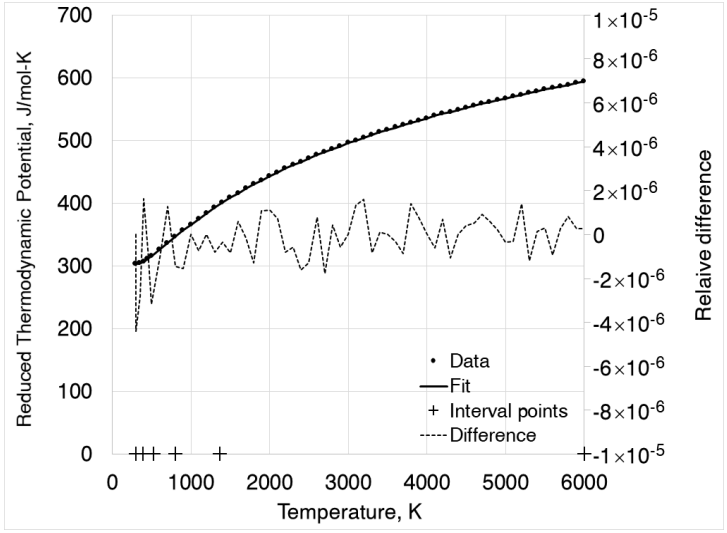
Sulfur Fluoride, Ion (SF_6^-) – F6S1-(g)

MW = 146.05 g/mol – $\Delta_f H^o$(298.15 K) = -1341.876 kJ/mol – S^o(298.15 K) = 302.86 J/mol-K – p^o = 0.1 MPa



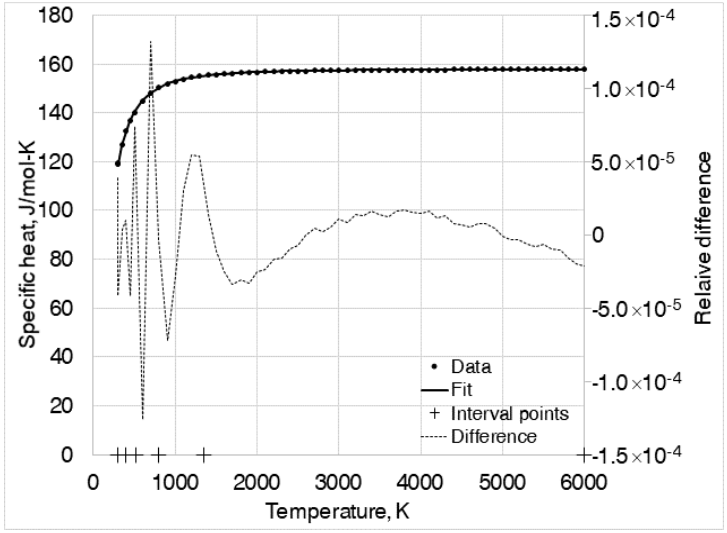


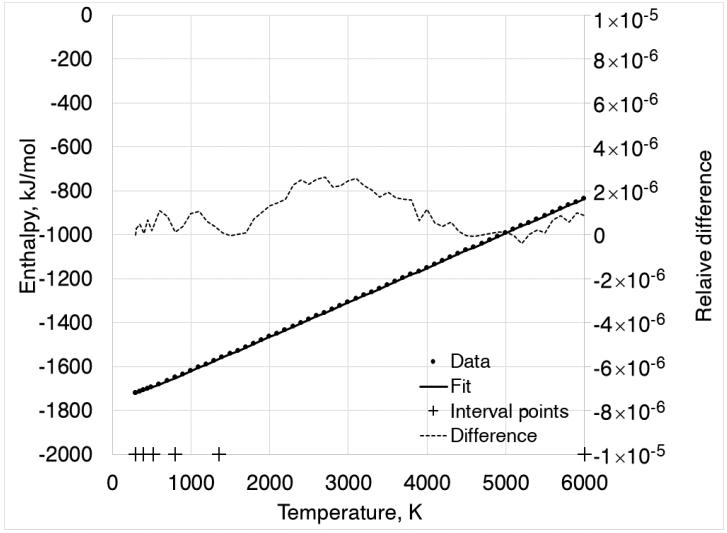


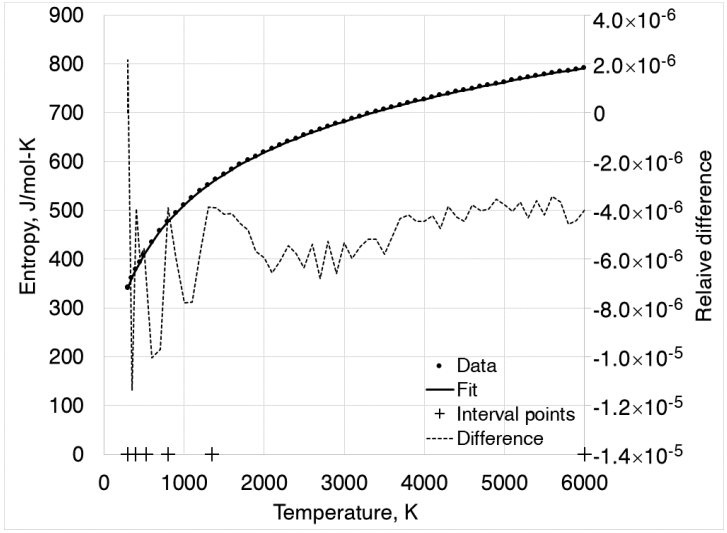


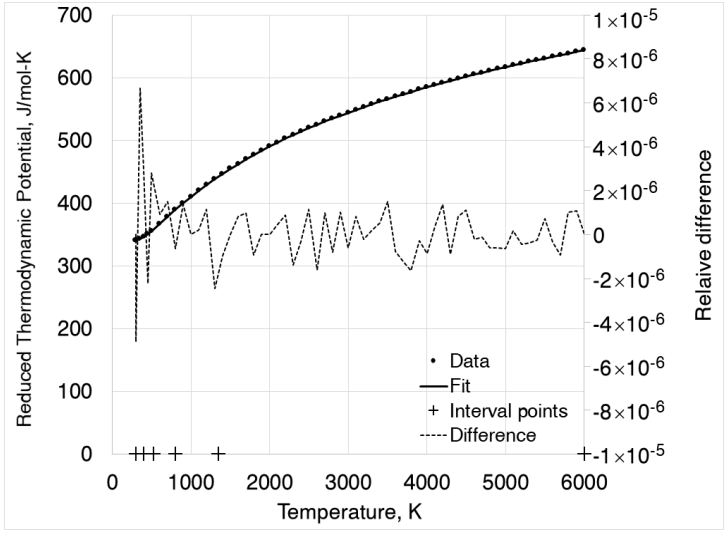
Tungsten Fluoride (WF_6) – F6W1(g)

MW = 297.84 g/mol – $\Delta_f H^o$(298.15 K) = -1721.716 kJ/mol – S^o(298.15 K) = 341.127 J/mol-K – p^o = 0.1 MPa



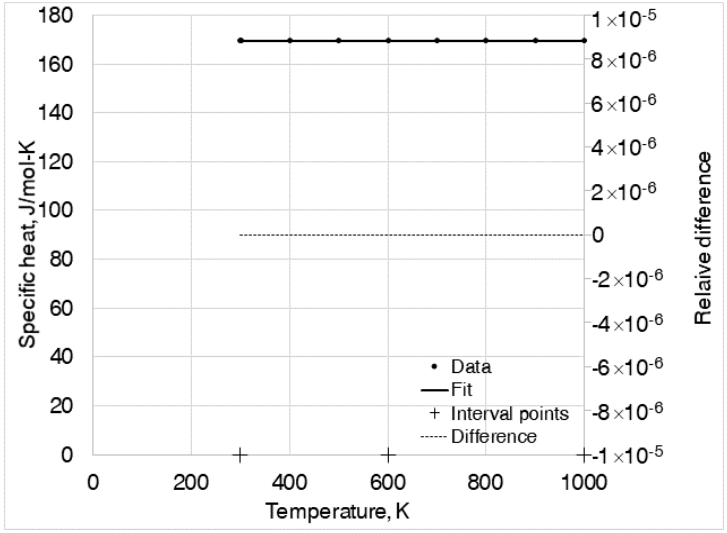


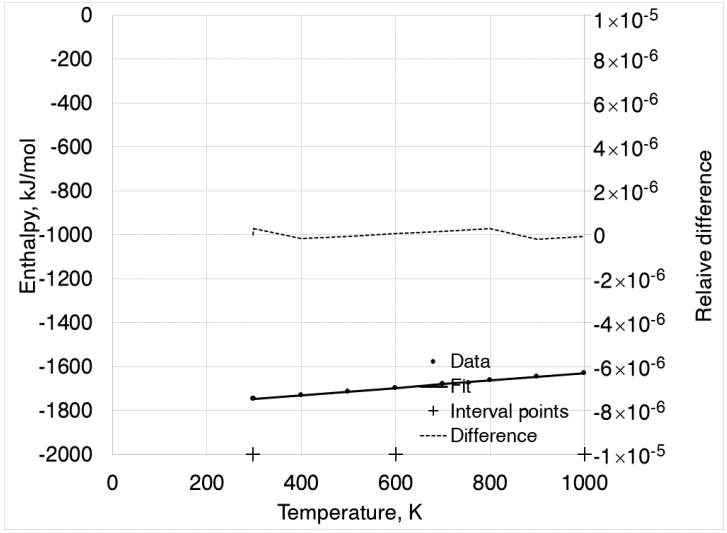


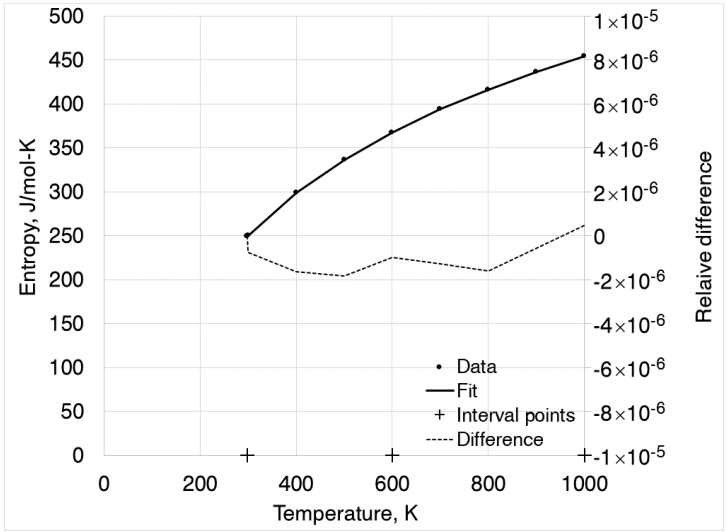


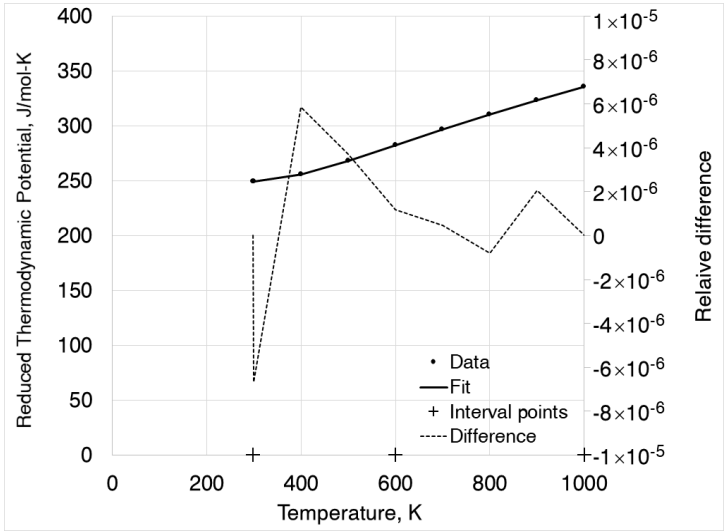
Tungsten Fluoride (WF_6) – F6W1(l)

MW = 297.84 g/mol – $\Delta_f H^o$(298.15 K) = -1748.368 kJ/mol – S^o(298.15 K) = 249.216 J/mol-K – p^o = 0.1 MPa



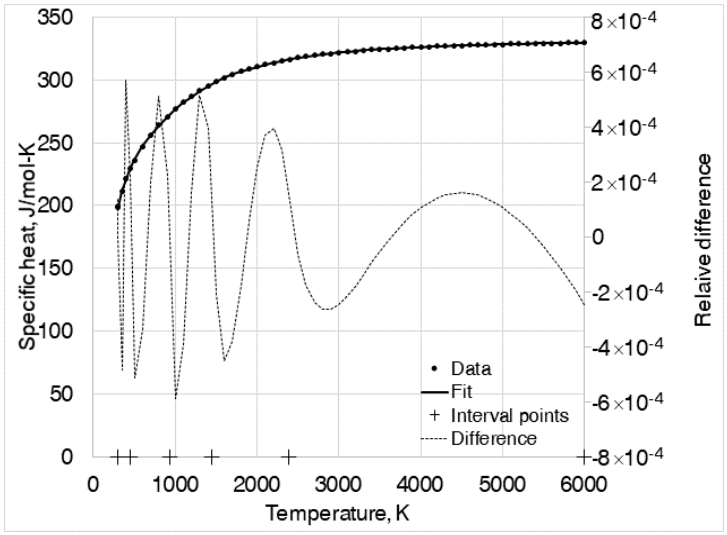


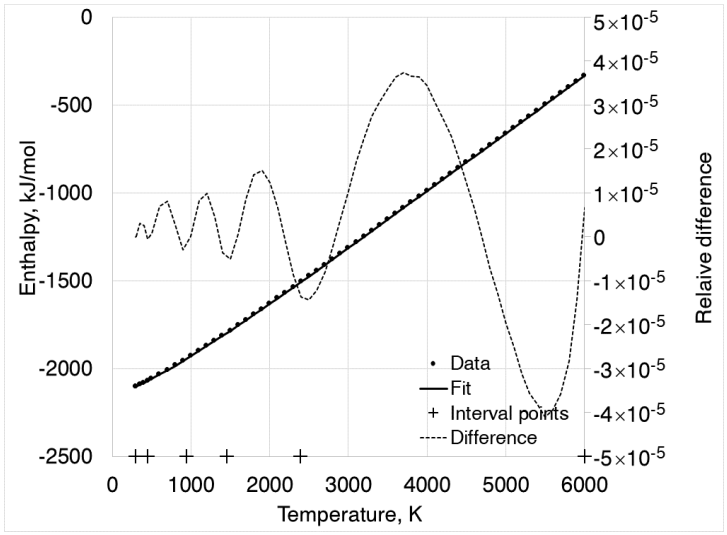


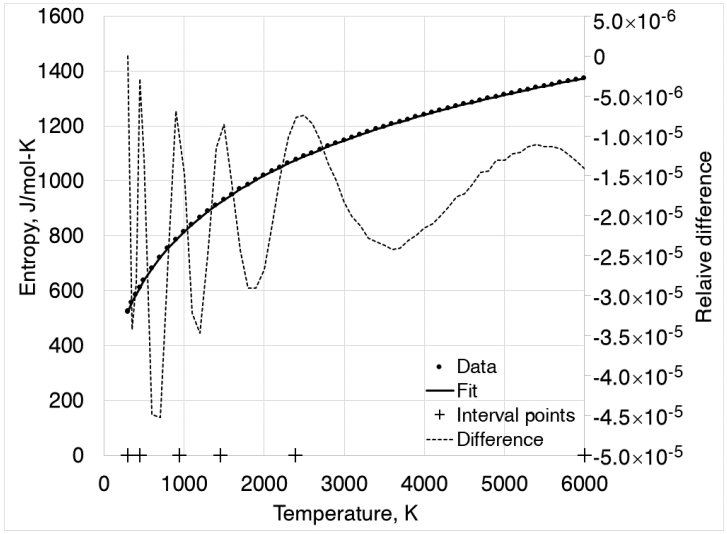


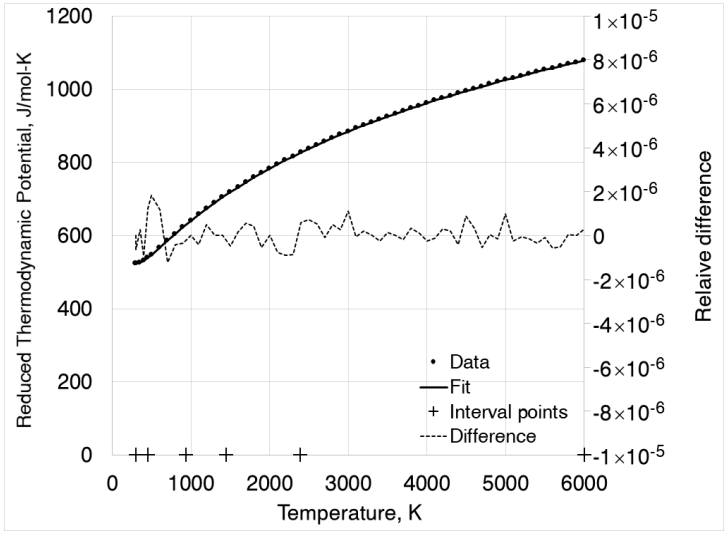
Hydrogen Fluoride (H_7F_7) – F7H7(g)

MW = 140.0444 g/mol – $\Delta_f H^o$(298.15 K) = -2102.544 kJ/mol – S^o(298.15 K) = 523.432 J/mol-K – p^o = 0.1 MPa



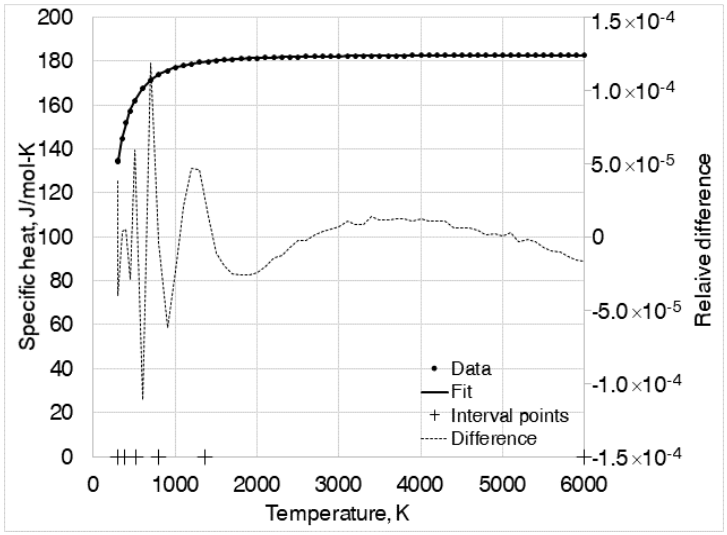


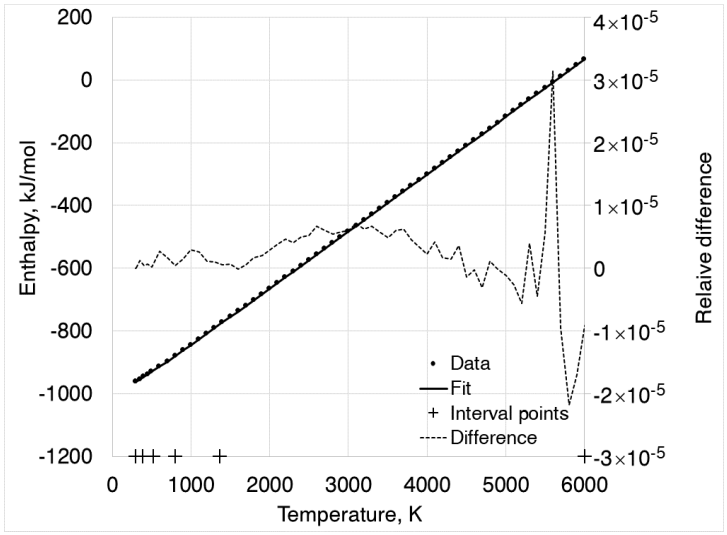


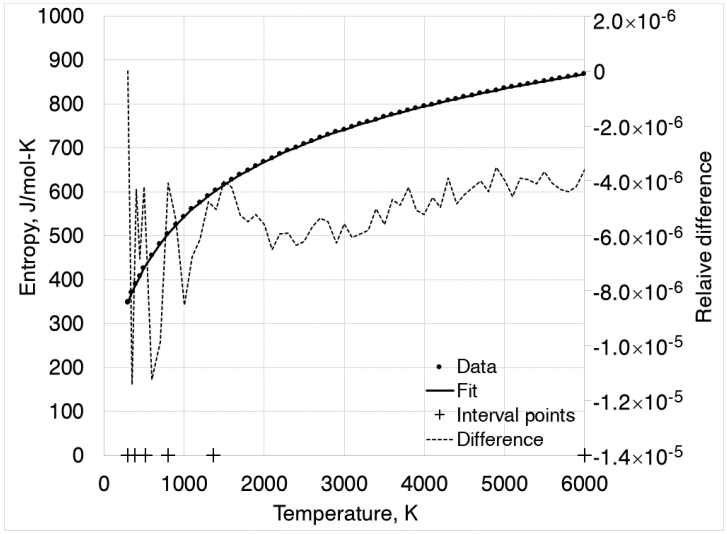


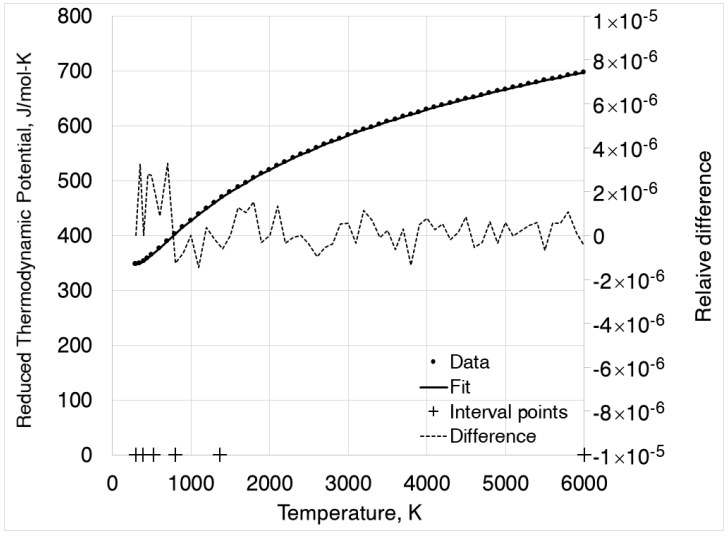
Iodine Fluoride (IF_7) – F7I1(g)

MW = 259.8933 g/mol – $\Delta_f H^o$(298.15 K) = -961.065 kJ/mol – S^o(298.15 K) = 347.74 J/mol-K – p^o = 0.1 MPa



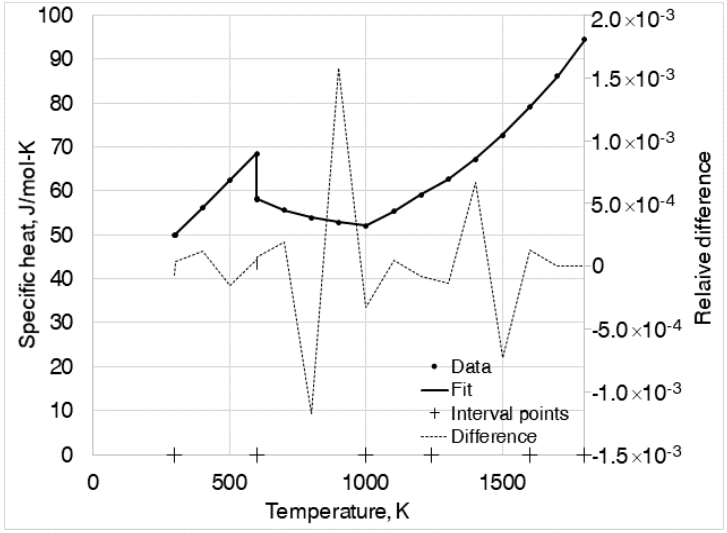


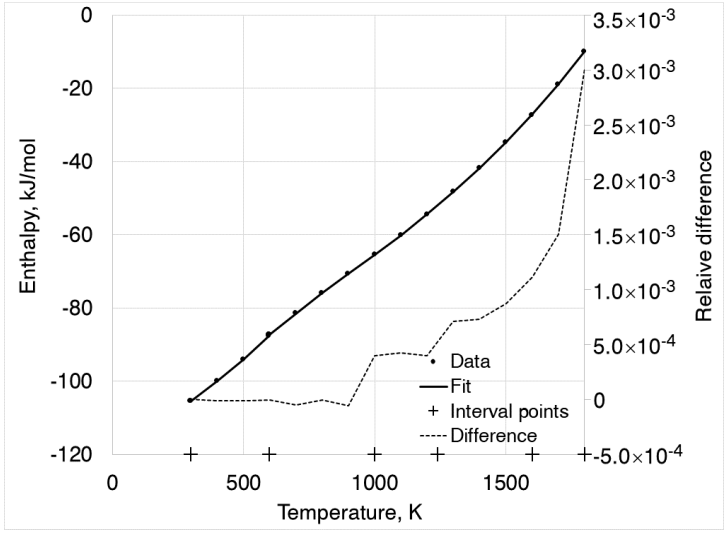


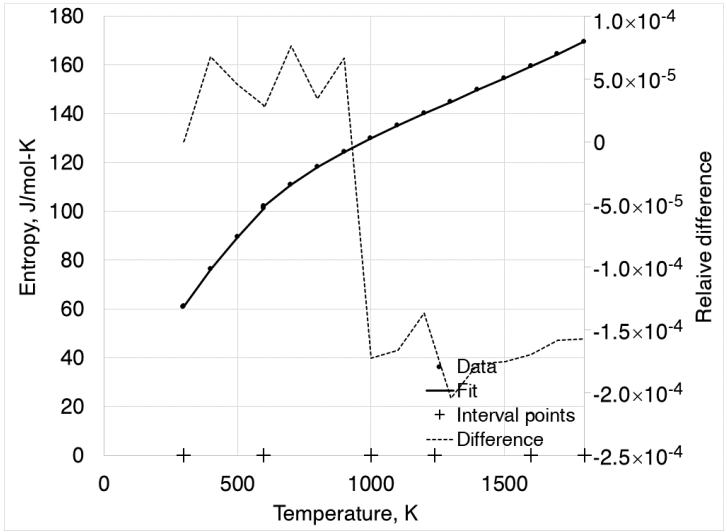


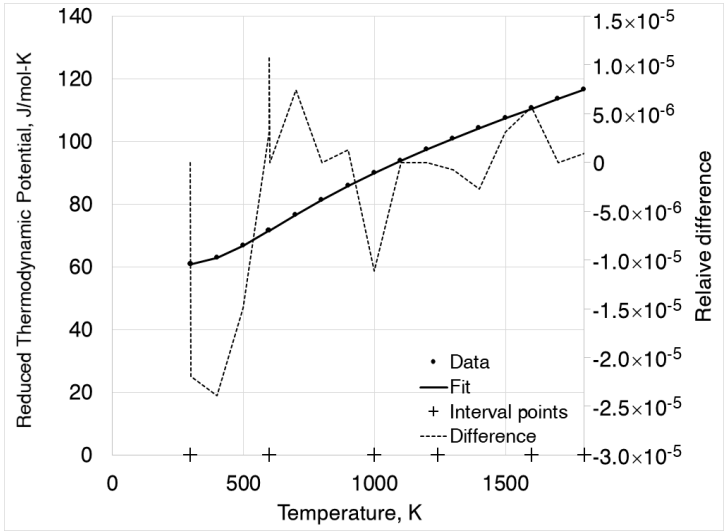
Iron Sulfide, Pyrrhotite ($Fe_{0.877}S$) – Fe0.877S1(cr)

MW = 81.04 g/mol – $\Delta_f H^o$(298.15 K) = -105.437 kJ/mol – S^o(298.15 K) = 60.799 J/mol-K – p^o = 0.1 MPa



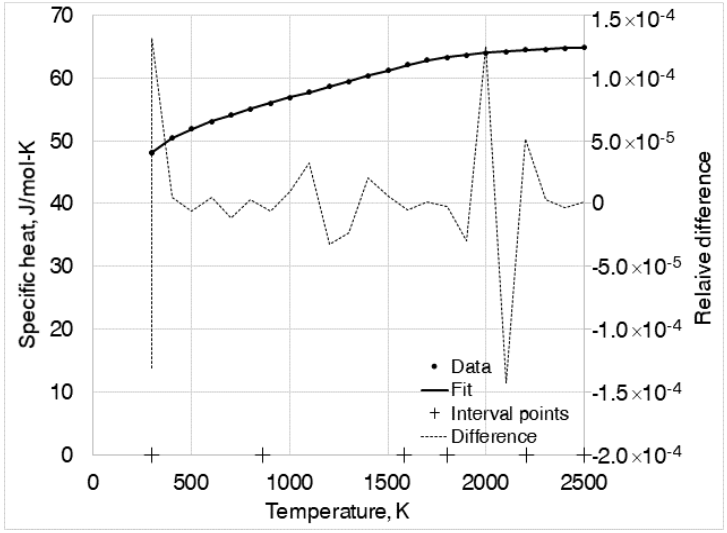


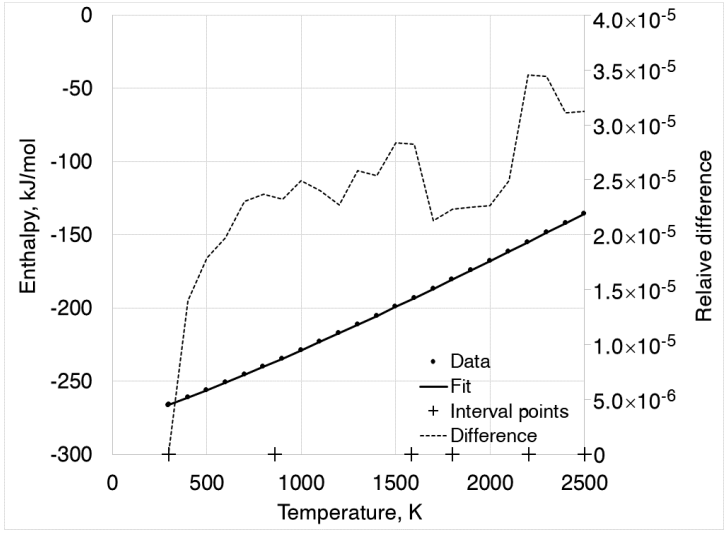


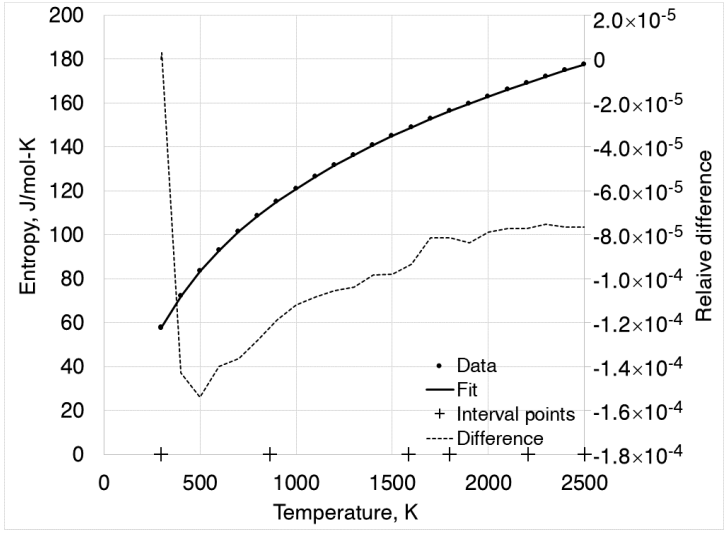


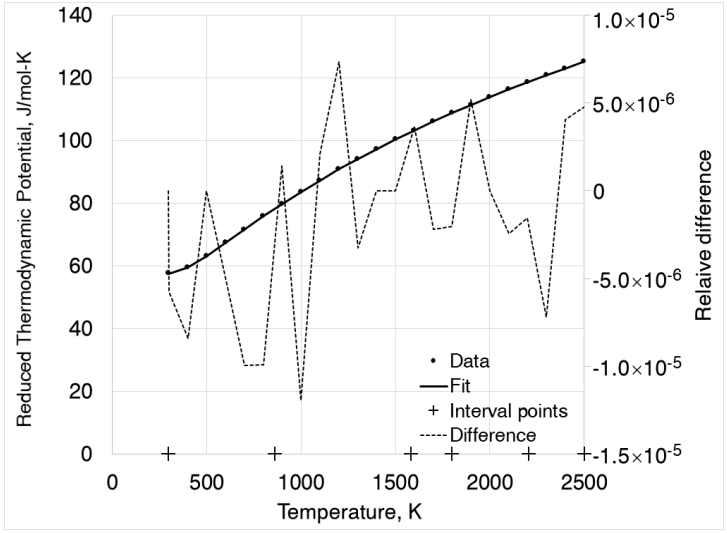
Iron Oxide, Wustite ($Fe_{0.947}O$) – Fe0.947O1(cr)

MW = 68.887 g/mol – $\Delta_f H^o$(298.15 K) = -266.27 kJ/mol – S^o(298.15 K) = 57.588 J/mol-K – p^o = 0.1 MPa



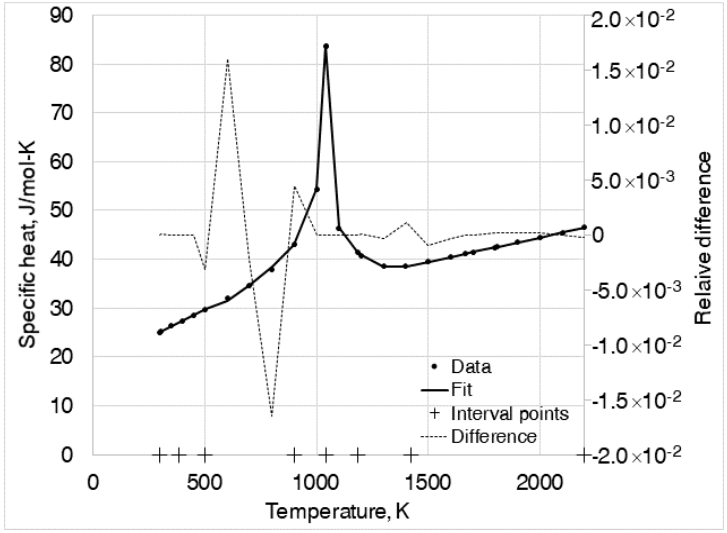


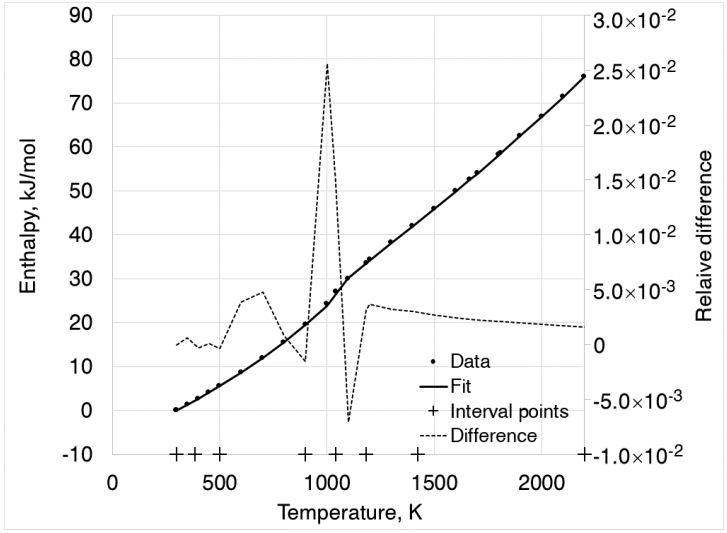


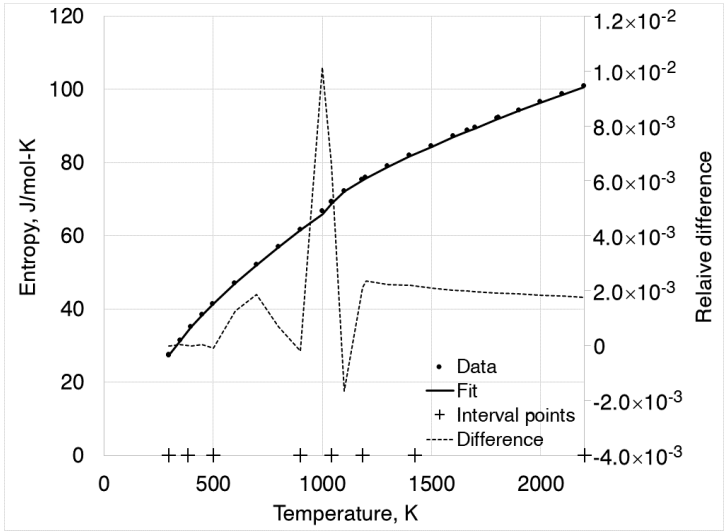


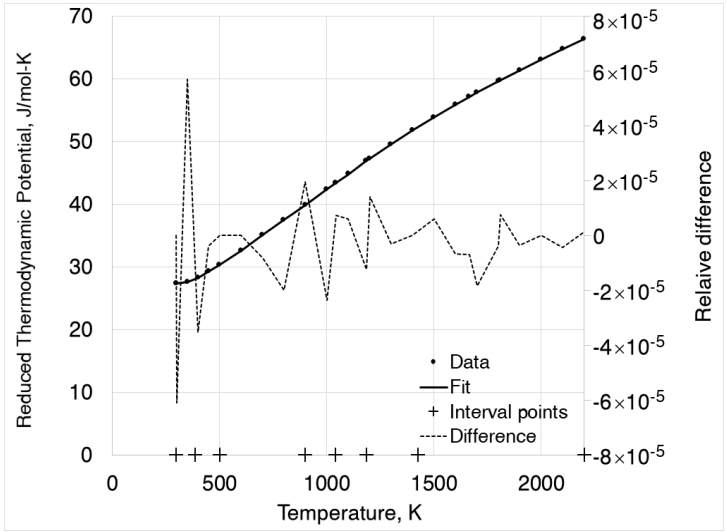
Iron, Alpha˙Delta (Fe) – Fe1(cr)_01

MW = 55.847 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 27.321 J/mol-K – p^o = 0.1 MPa



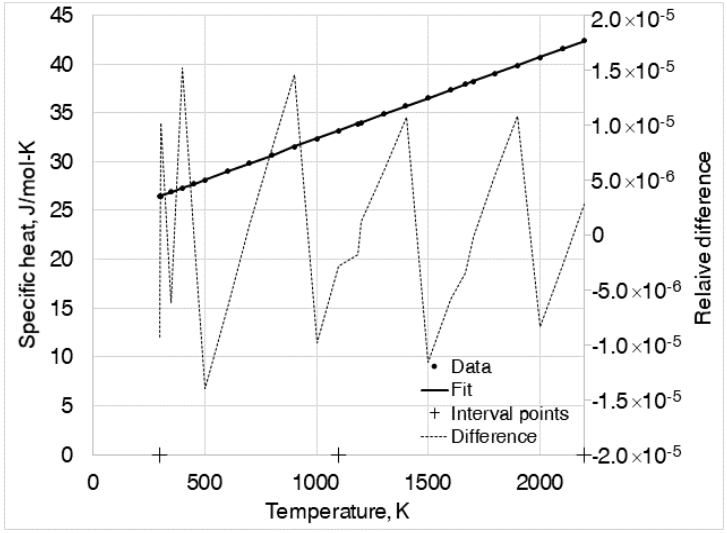


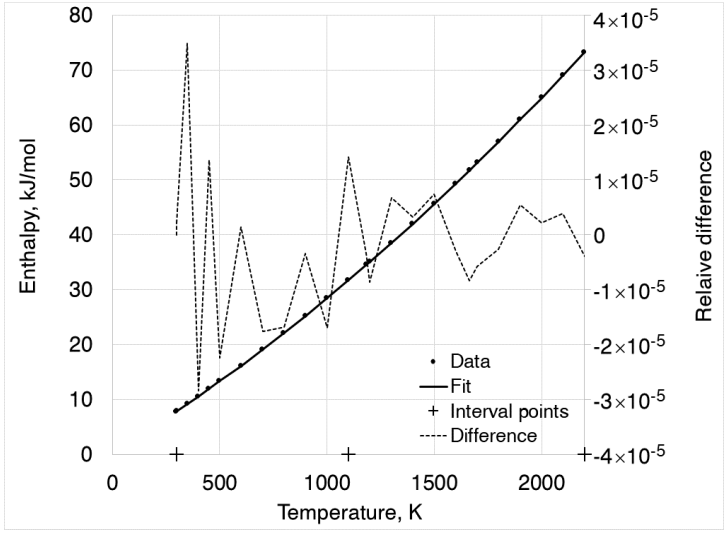


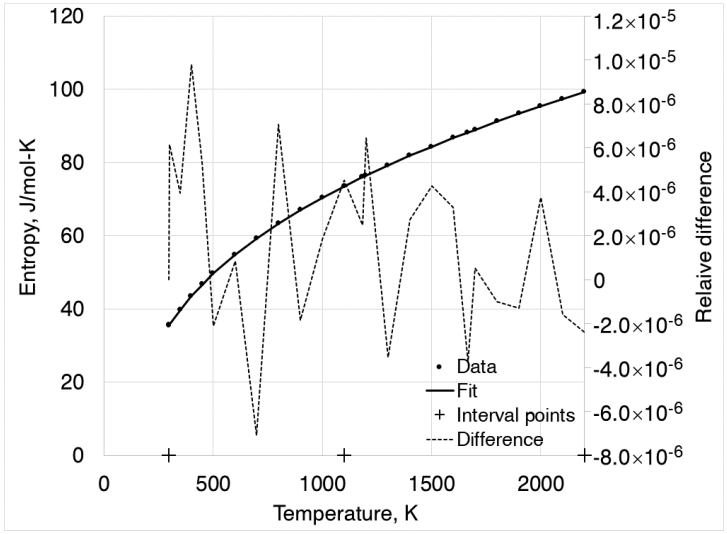


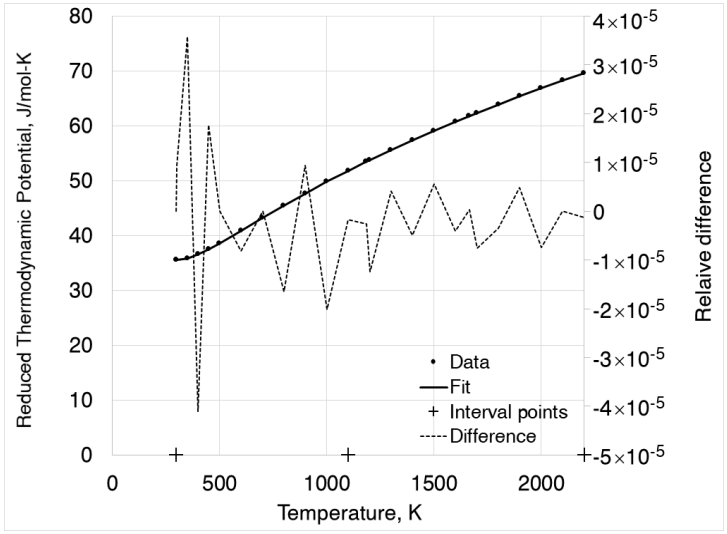
Iron, Gamma (Fe) – Fe1(cr)_02

MW = 55.847 g/mol – $\Delta_f H^o$(298.15 K) = 7.788 kJ/mol – S^o(298.15 K) = 35.545 J/mol-K – p^o = 0.1 MPa



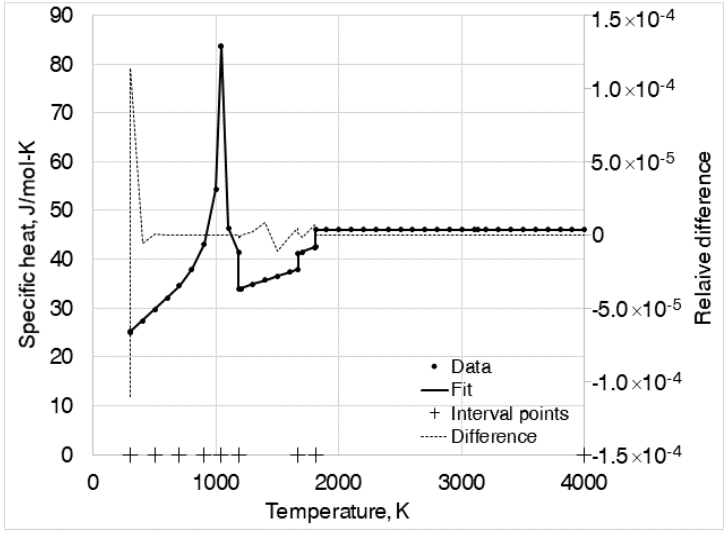


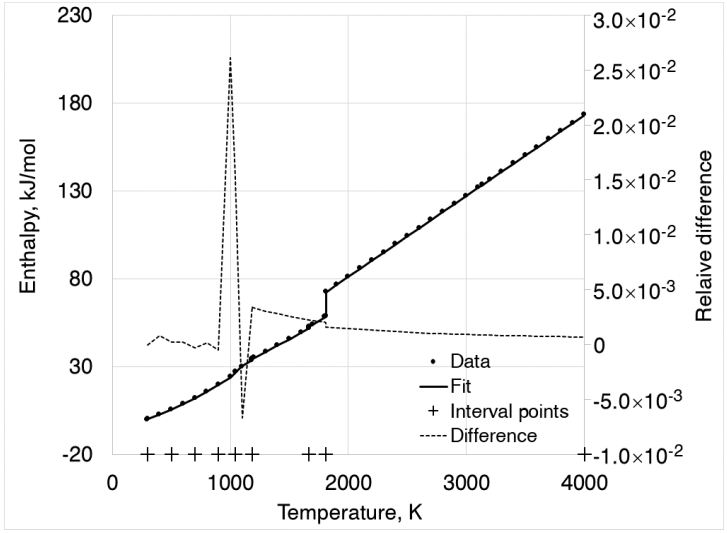


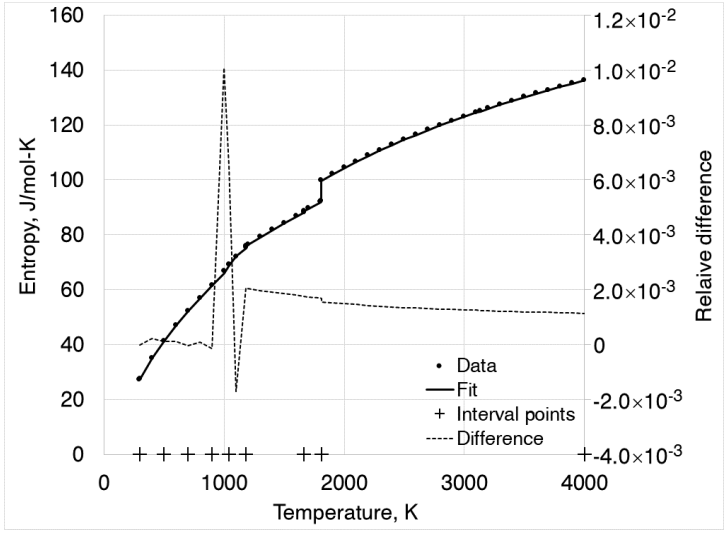


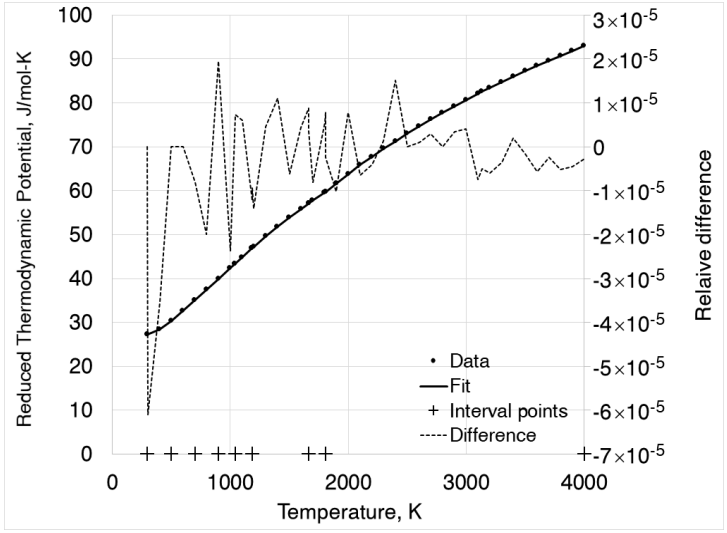
Iron (Fe) – Fe1(cr,l)

MW = 55.847 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 27.321 J/mol-K – p^o = 0.1 MPa



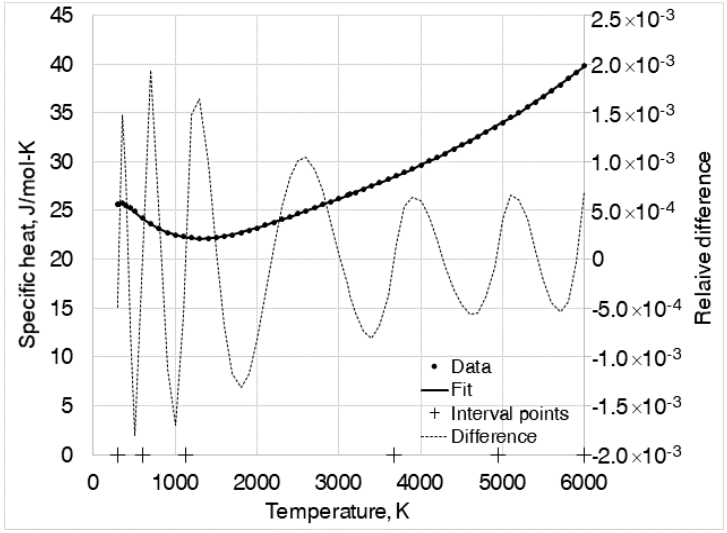


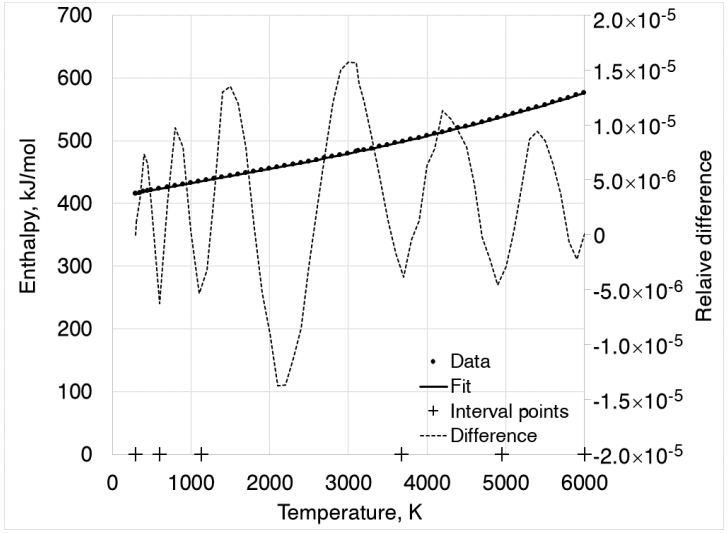


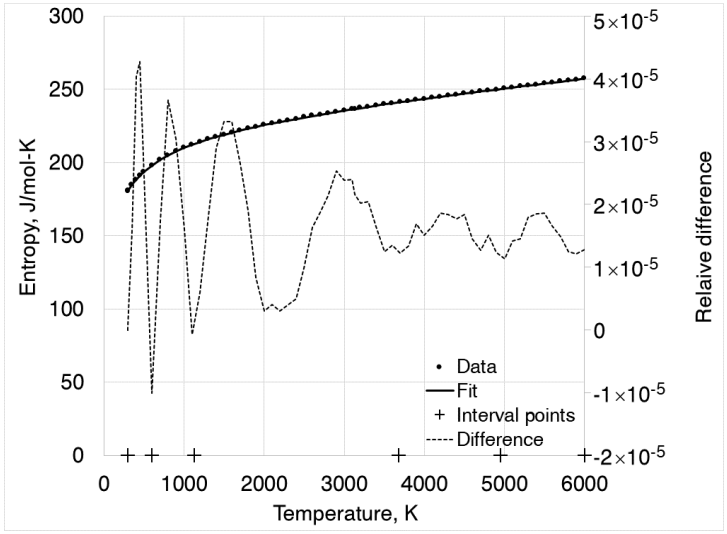


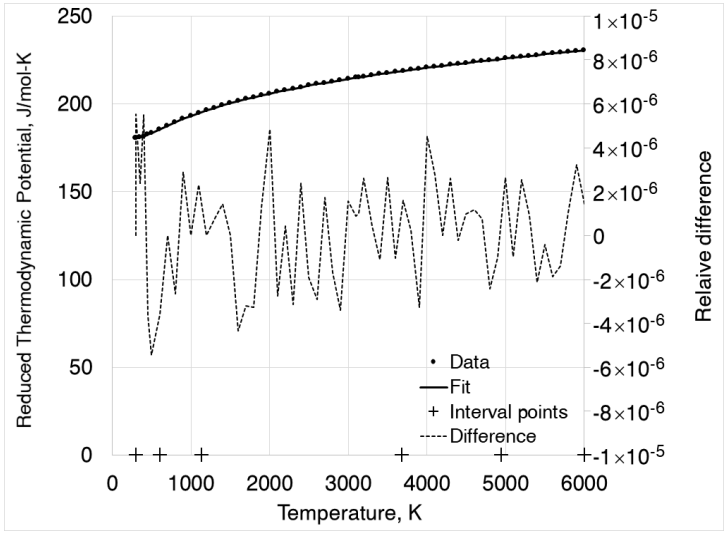
Iron (Fe) – Fe1(g)

MW = 55.847 g/mol – $\Delta_f H^o$(298.15 K) = 415.471 kJ/mol – S^o(298.15 K) = 180.488 J/mol-K – p^o = 0.1 MPa



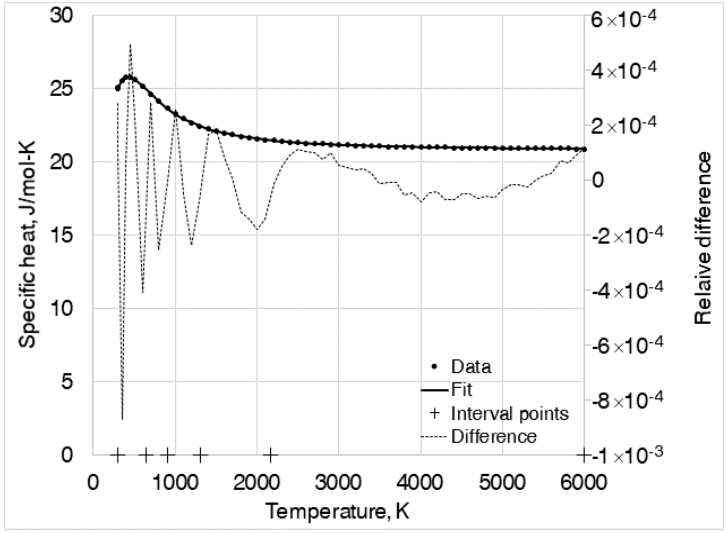


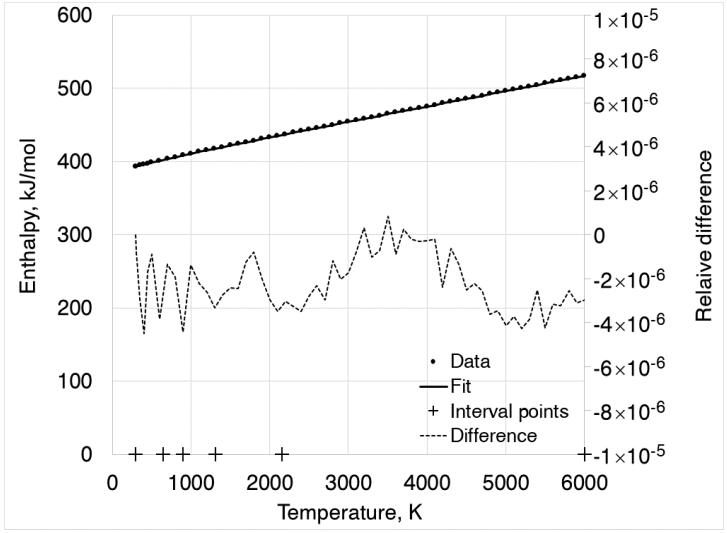


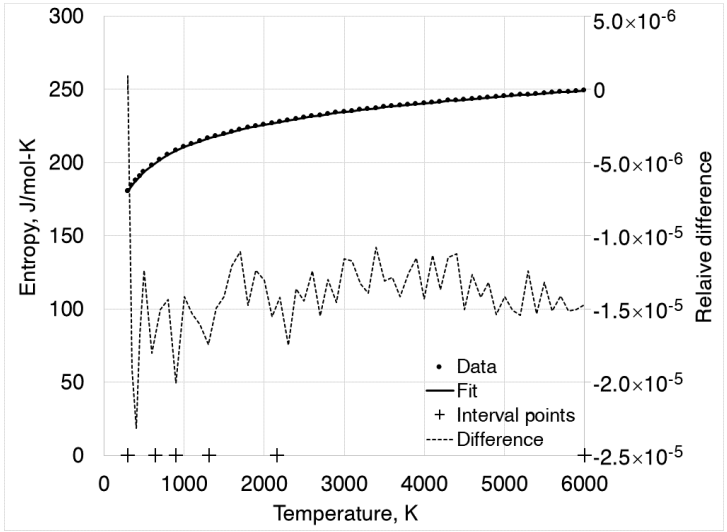


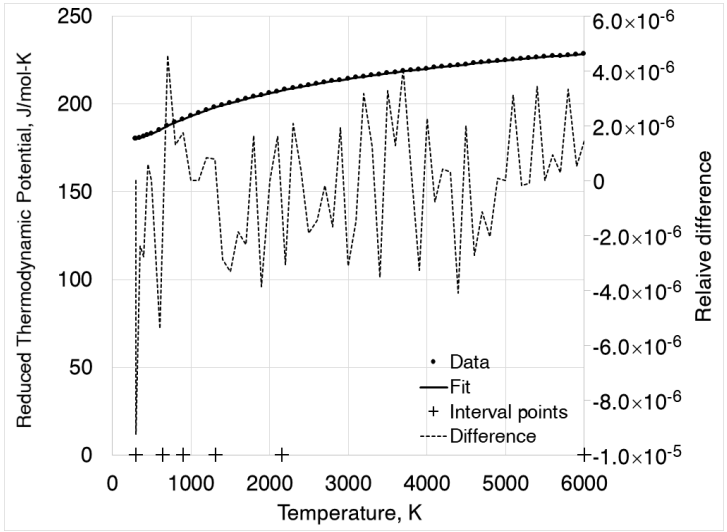
Iron, Ion (Fe^-) – Fe1-(g)

MW = 55.848 g/mol – $\Delta_f H^o$(298.15 K) = 393.338 kJ/mol – S^o(298.15 K) = 180.198 J/mol-K – p^o = 0.1 MPa



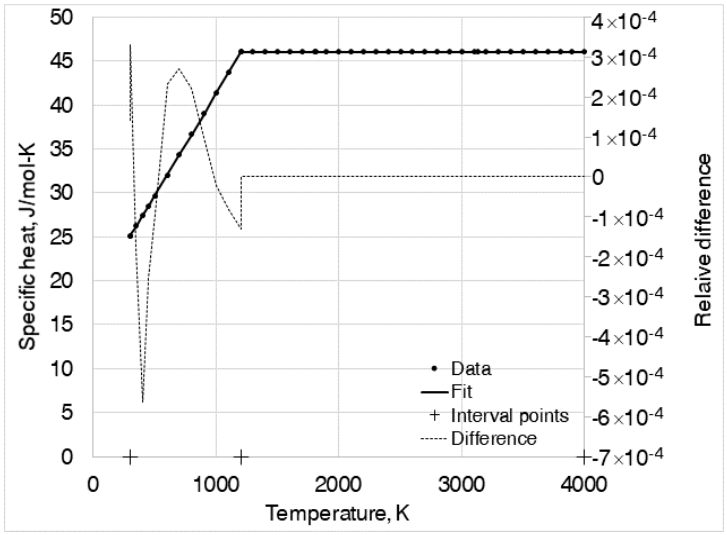


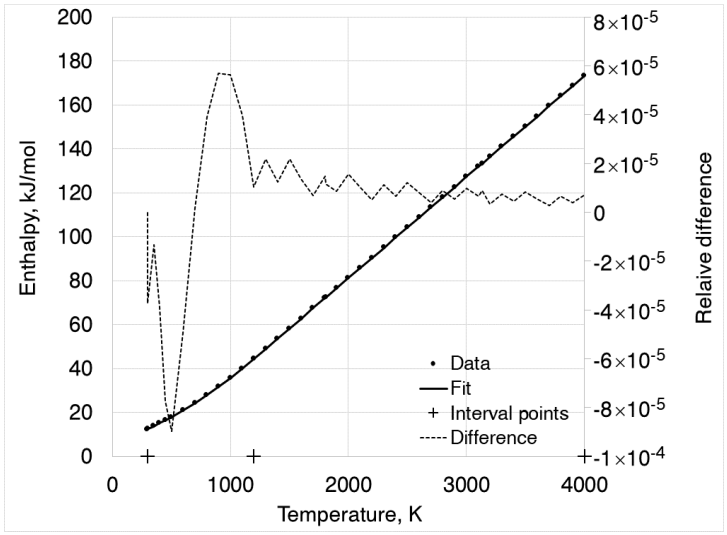


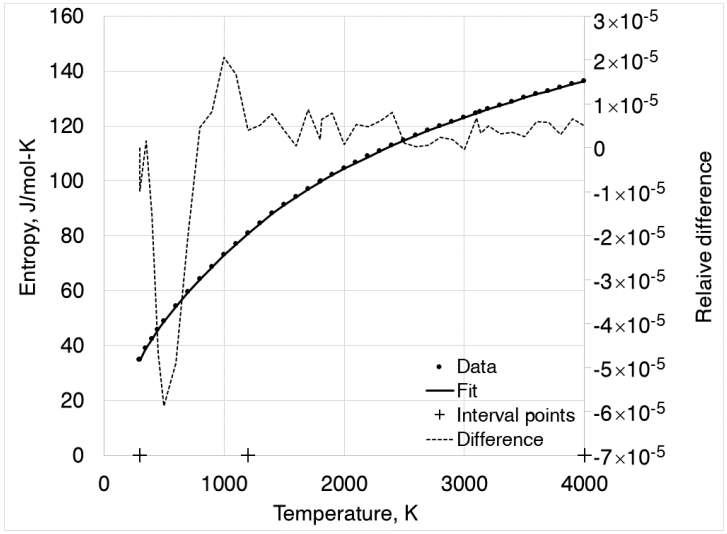


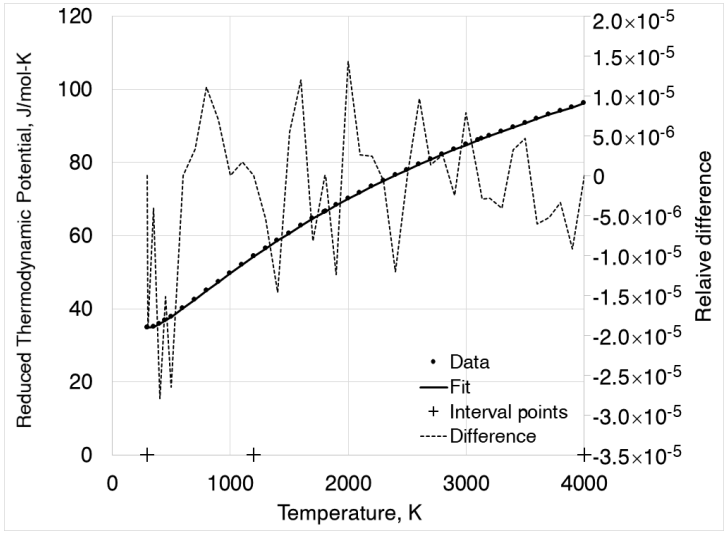
Iron (Fe) – Fe1(l)

MW = 55.847 g/mol – $\Delta_f H^o$(298.15 K) = 12.395 kJ/mol – S^o(298.15 K) = 34.763 J/mol-K – p^o = 0.1 MPa



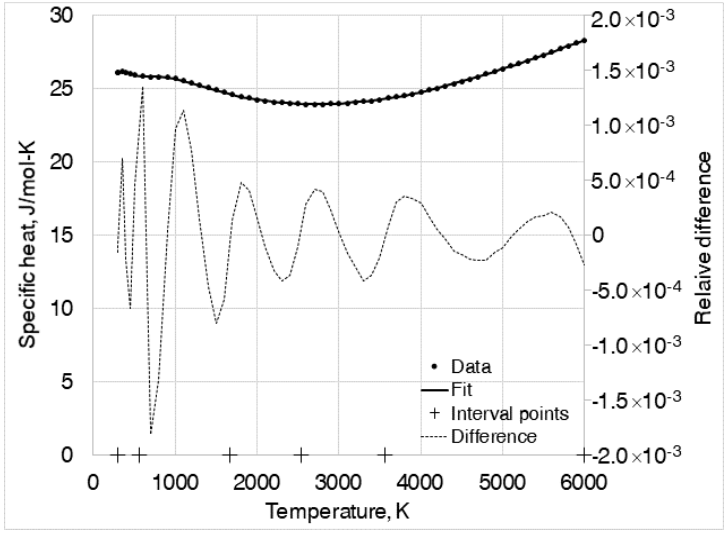


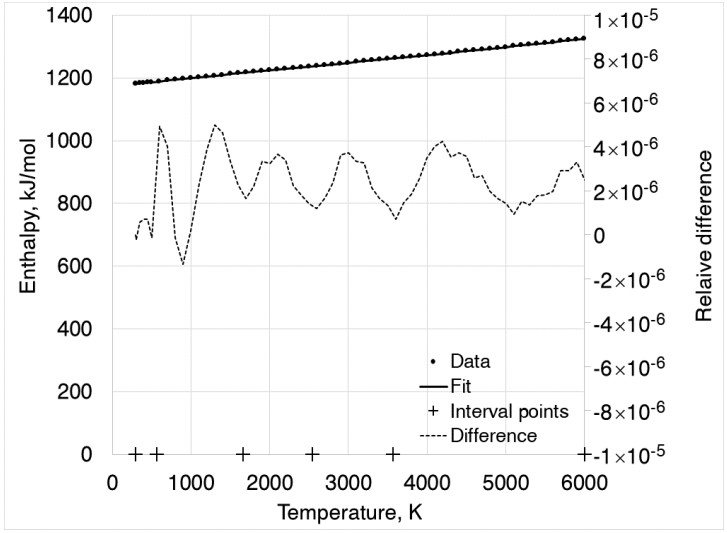


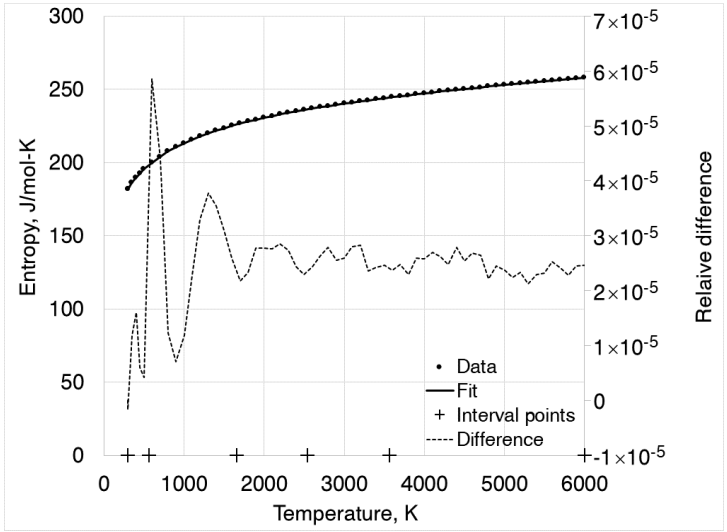


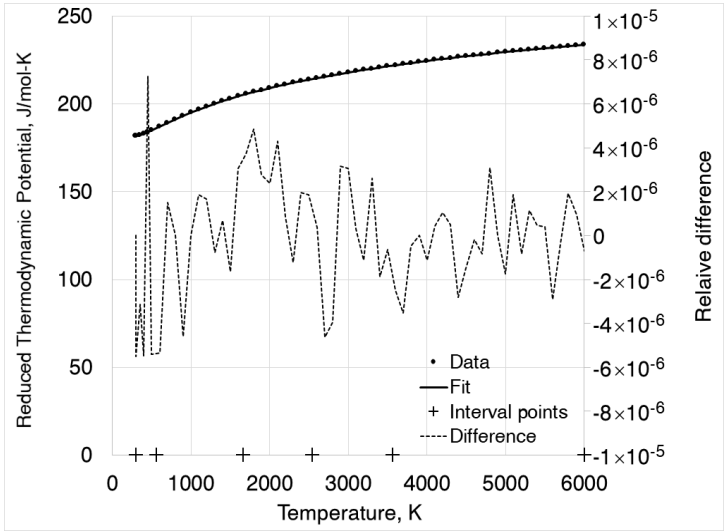
Iron, Ion (Fe^+) – Fe1+(g)

MW = 55.846 g/mol – $\Delta_f H^o$(298.15 K) = 1181.144 kJ/mol – S^o(298.15 K) = 181.857 J/mol-K – p^o = 0.1 MPa



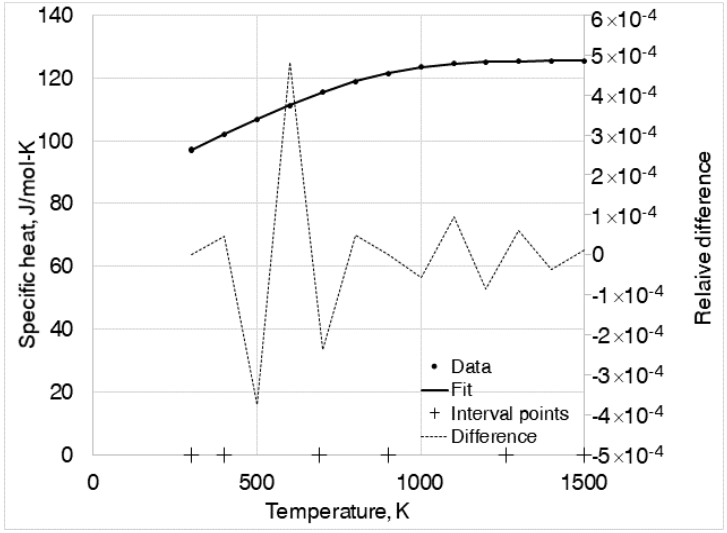


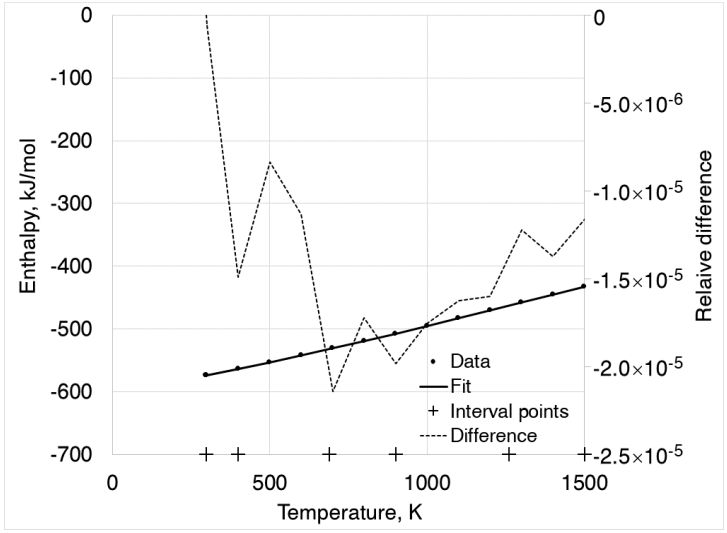


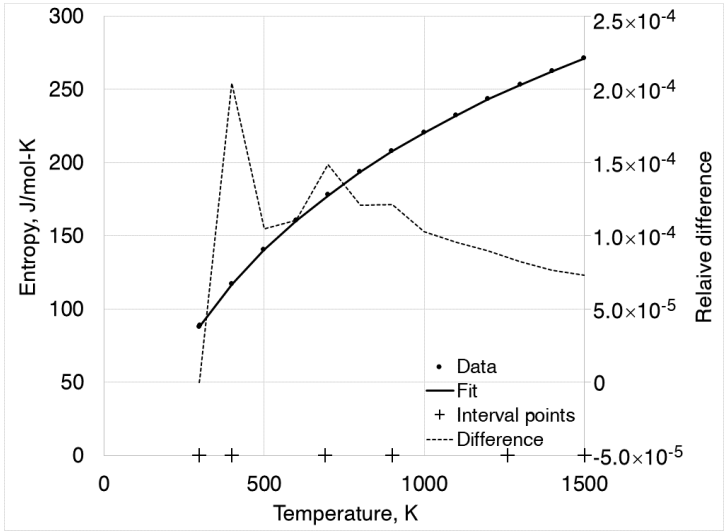


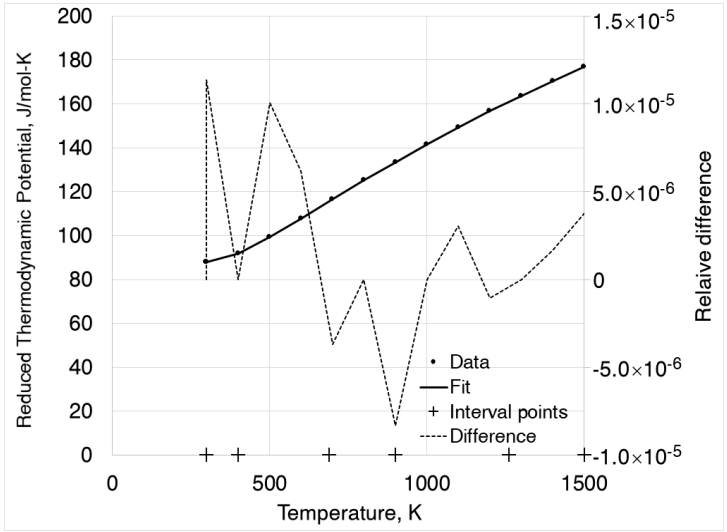
Iron Hydroxide ($Fe(OH)_2$) – Fe1H2O2(cr)

MW = 89.862 g/mol – $\Delta_f H^o$(298.15 K) = -574.045 kJ/mol – S^o(298.15 K) = 87.864 J/mol-K – p^o = 0.1 MPa



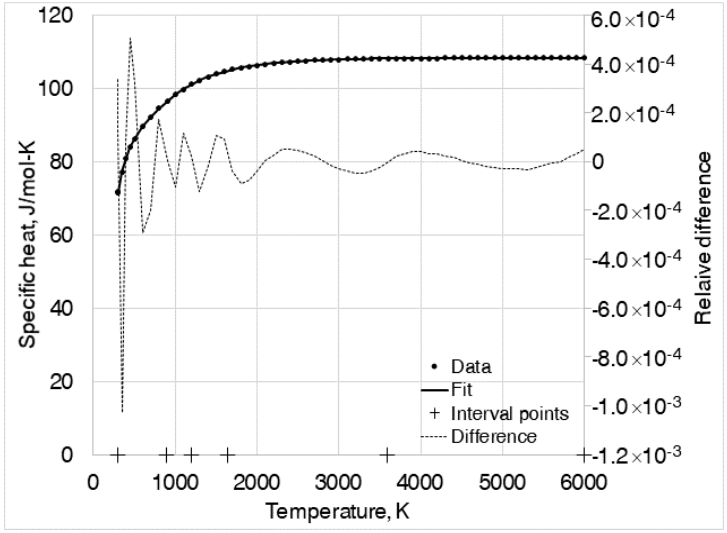


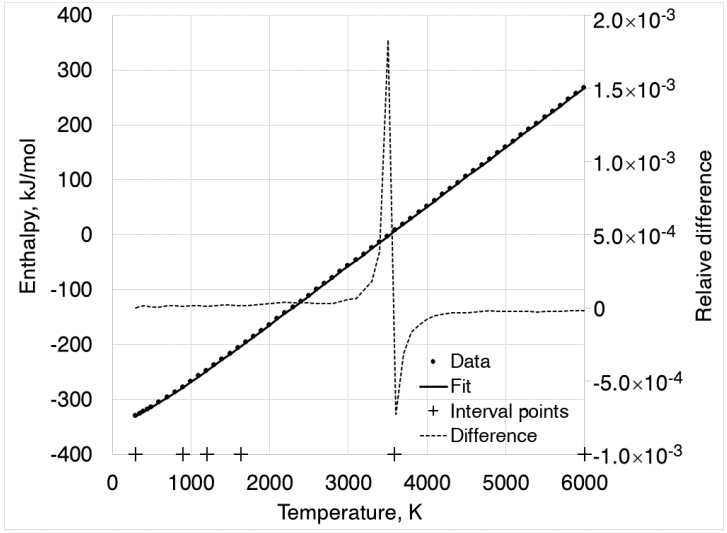


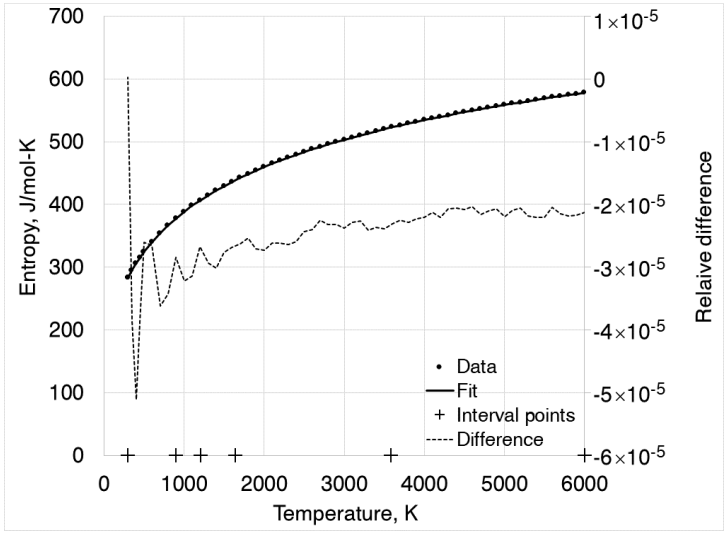


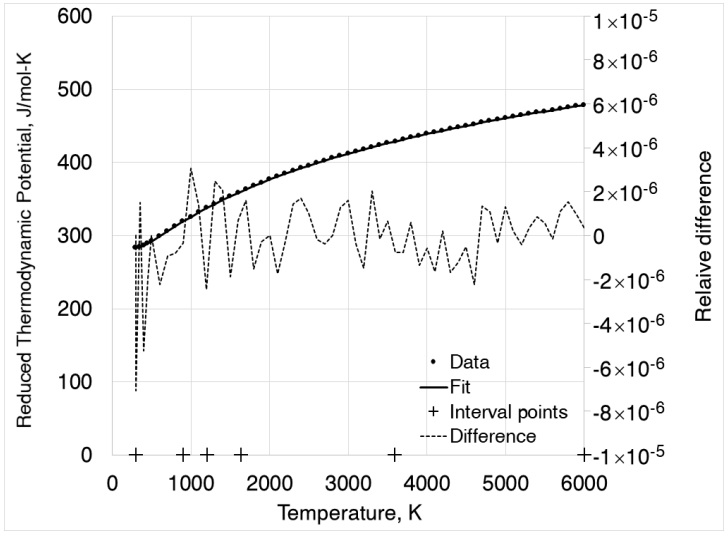
Iron Hydroxide ($Fe(OH)_2$) – Fe1H2O2(g)

MW = 89.862 g/mol – $\Delta_f H^o$(298.15 K) = -330.536 kJ/mol – S^o(298.15 K) = 283.088 J/mol-K – p^o = 0.1 MPa



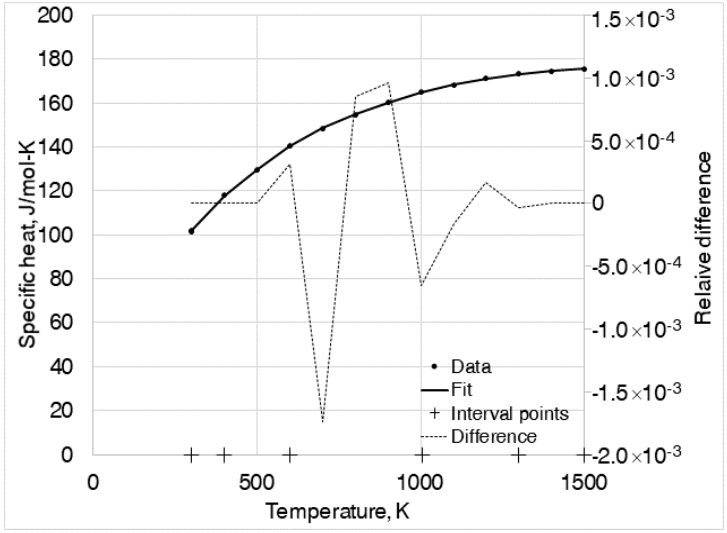


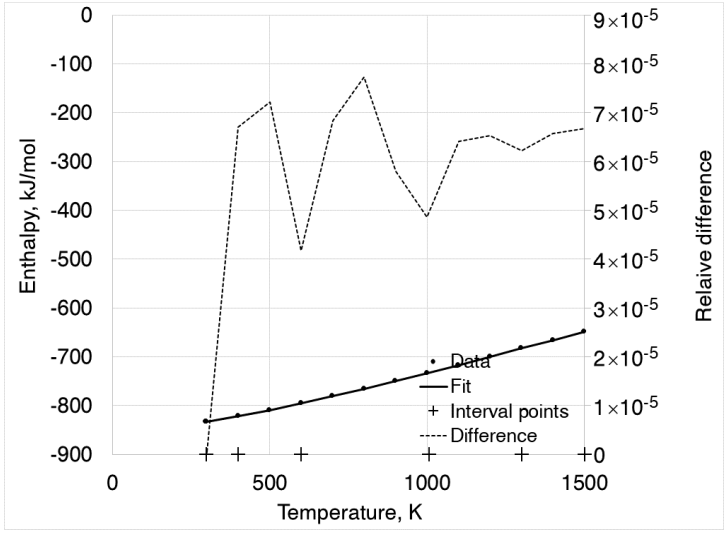


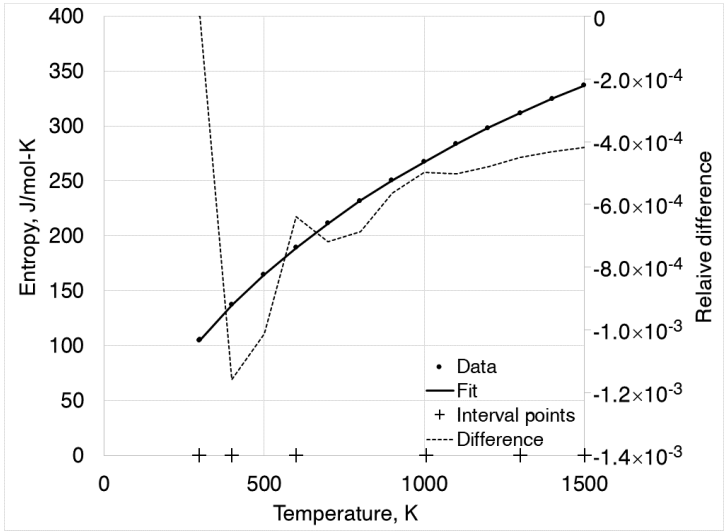


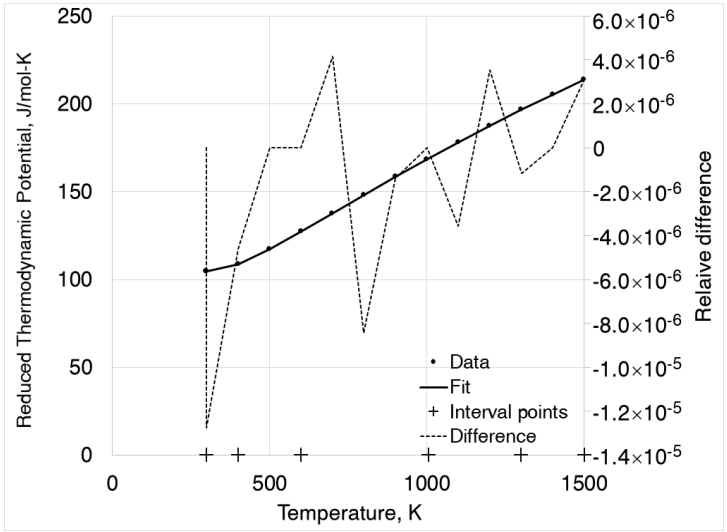
Iron Hydroxide ($Fe(OH)_3$) – Fe1H3O3(cr)

MW = 106.869 g/mol – $\Delta_f H^o$(298.15 K) = -832.616 kJ/mol – S^o(298.15 K) = 104.6 J/mol-K – p^o = 0.1 MPa



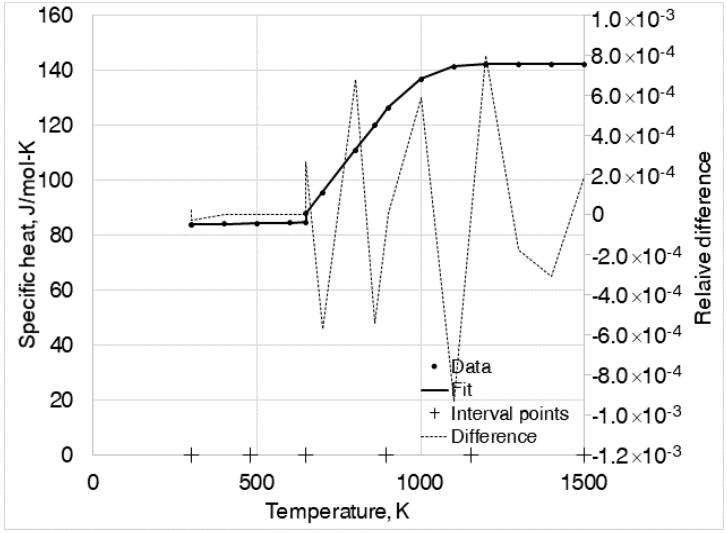


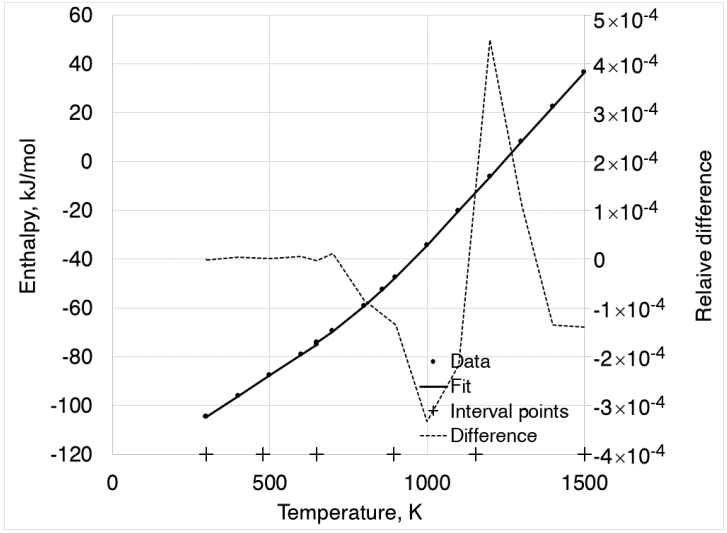


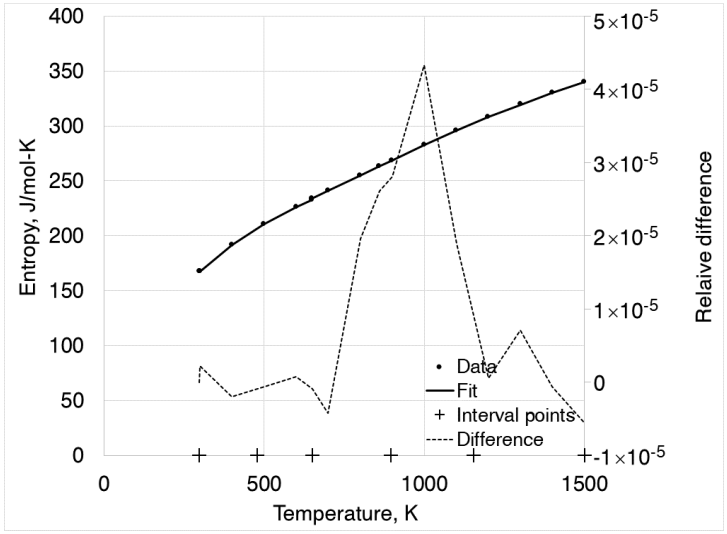


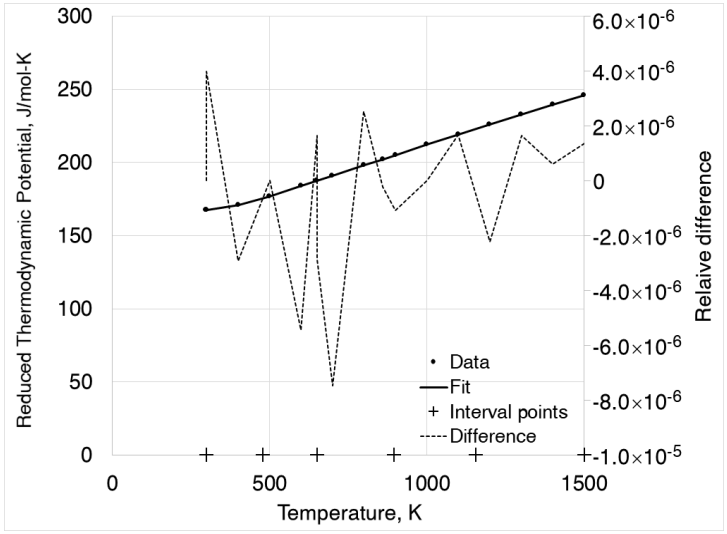
Iron Iodide (FeI_2) – Fe1I2(cr)

MW = 309.656 g/mol – $\Delta_f H^o$(298.15 K) = -104.6 kJ/mol – S^o(298.15 K) = 167.36 J/mol-K – p^o = 0.1 MPa



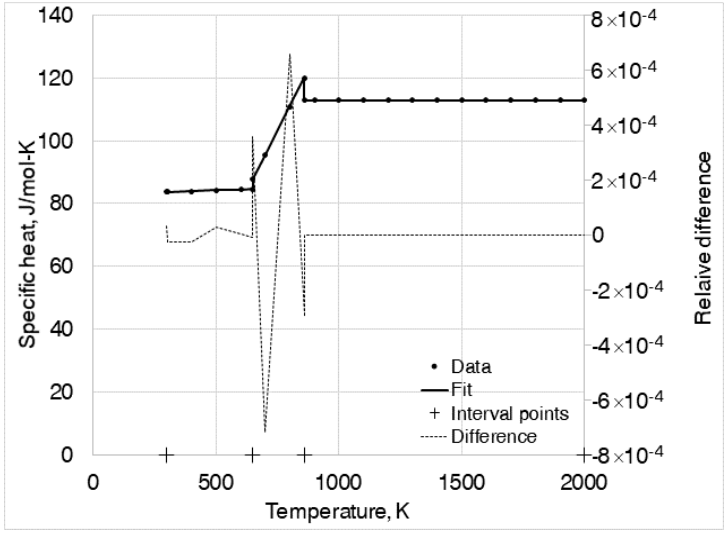


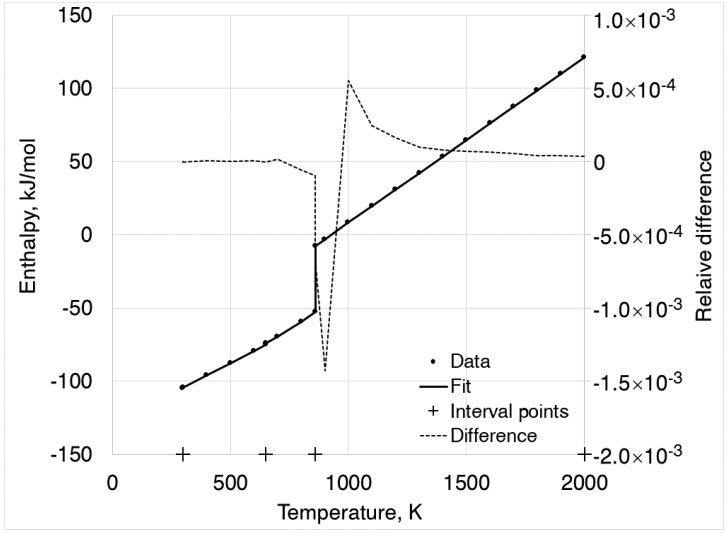


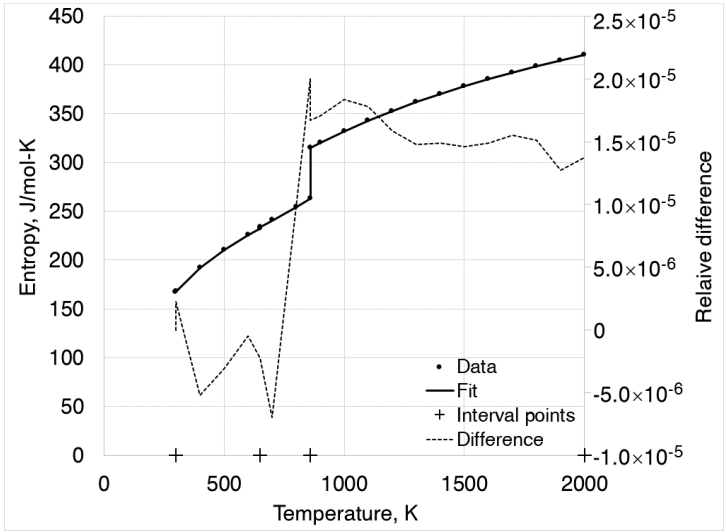


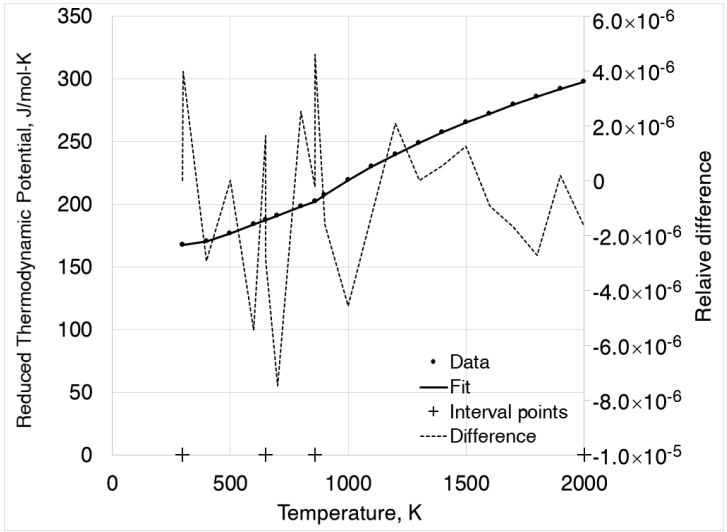
Iron Iodide (FeI_2) – Fe1I2(cr,l)

MW = 309.656 g/mol – $\Delta_f H^o$(298.15 K) = -104.6 kJ/mol – S^o(298.15 K) = 167.36 J/mol-K – p^o = 0.1 MPa



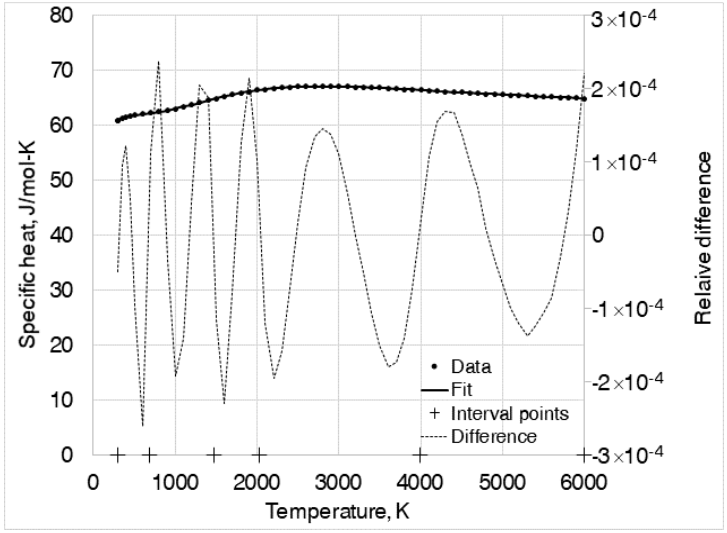


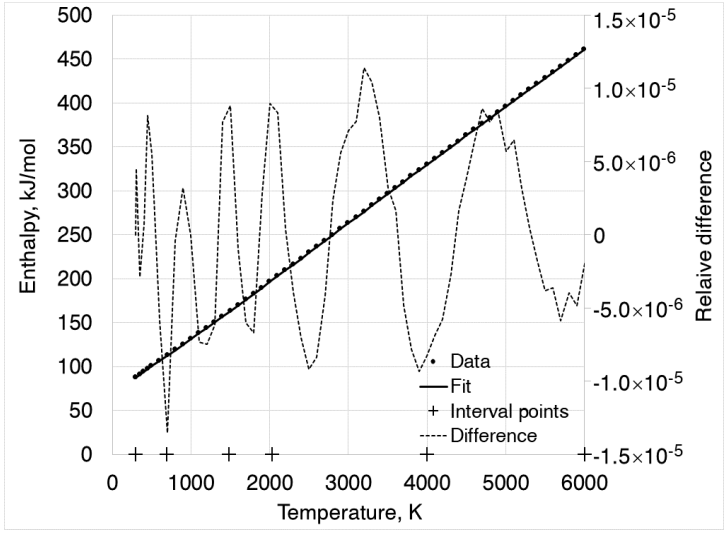


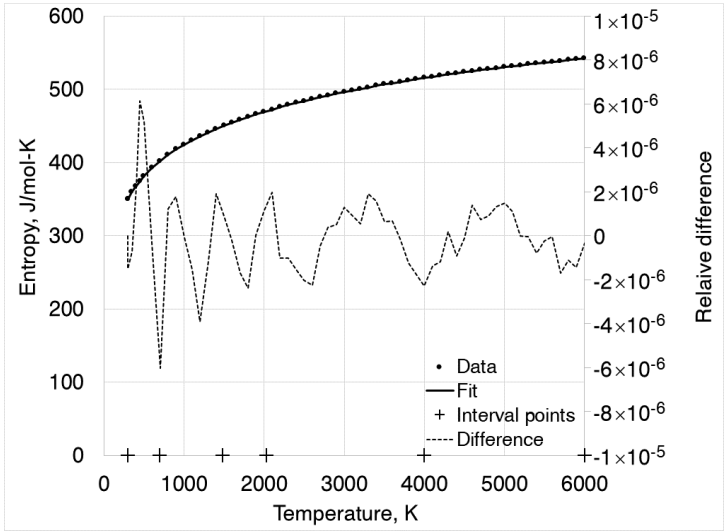


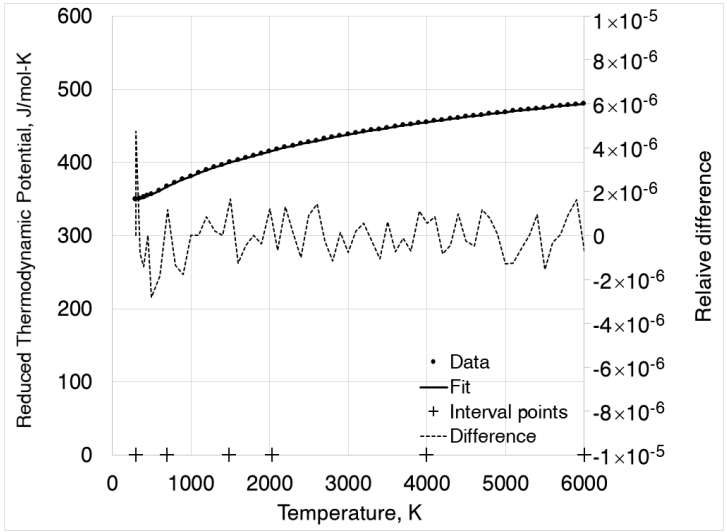
Iron Iodide (FeI_2) – Fe1I2(g)

MW = 309.656 g/mol – $\Delta_f H^o$(298.15 K) = 87.864 kJ/mol – S^o(298.15 K) = 349.597 J/mol-K – p^o = 0.1 MPa



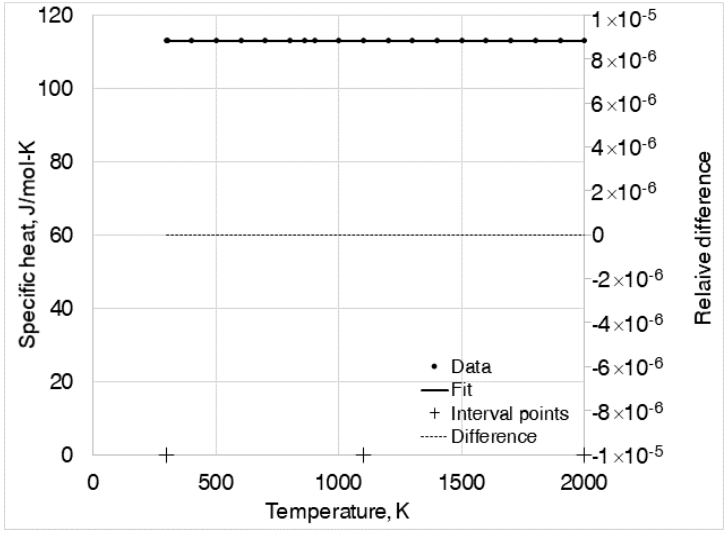


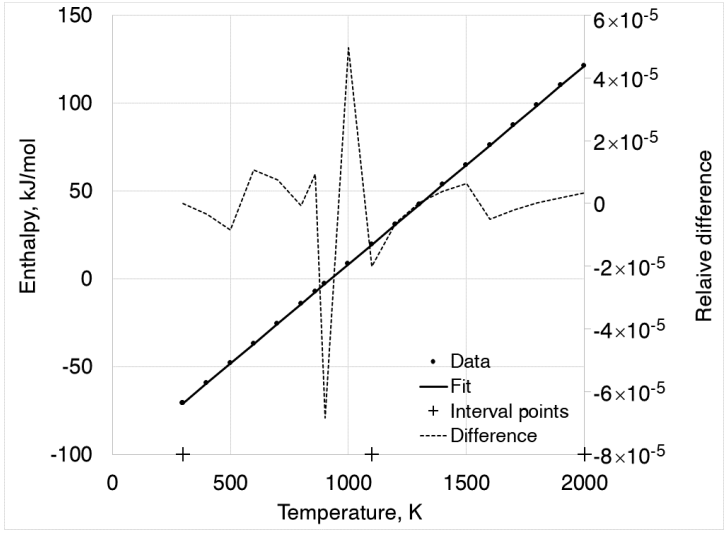


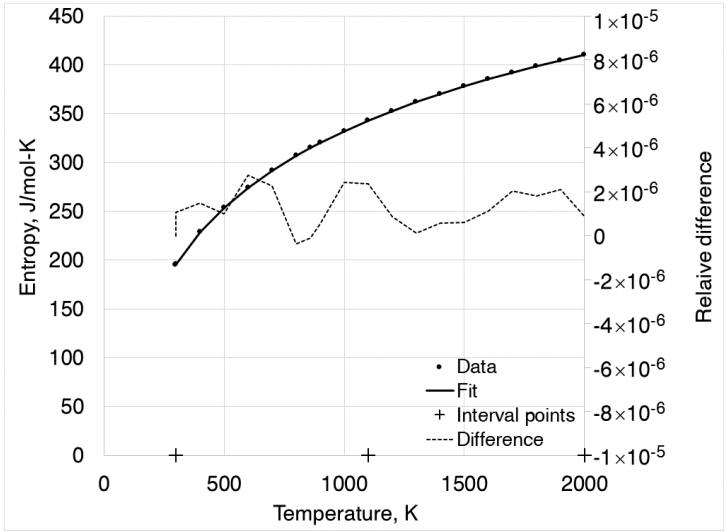


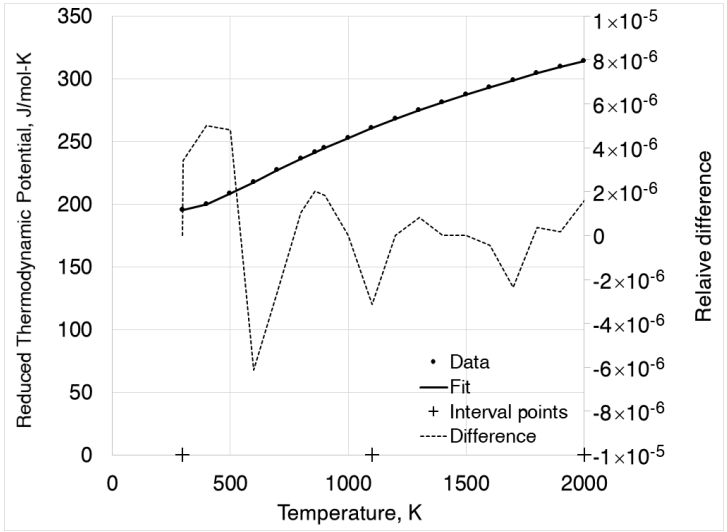
Iron Iodide (FeI_2) – Fe1I2(l)

MW = 309.656 g/mol – $\Delta_f H^o$(298.15 K) = -71.046 kJ/mol – S^o(298.15 K) = 195.43 J/mol-K – p^o = 0.1 MPa



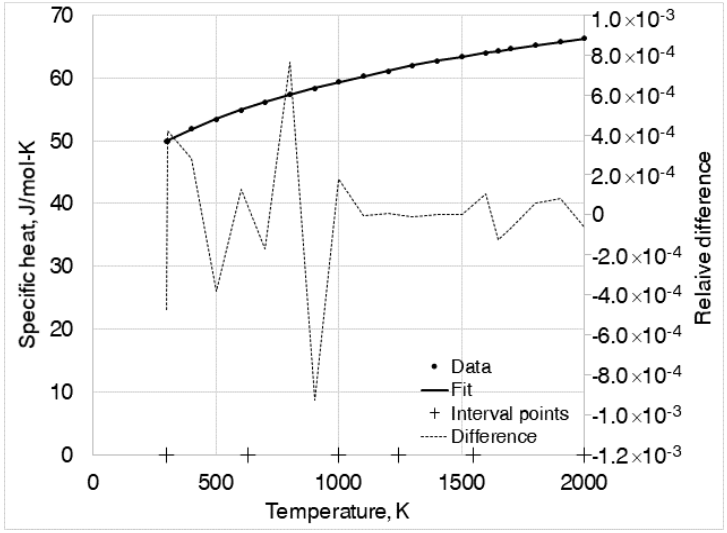


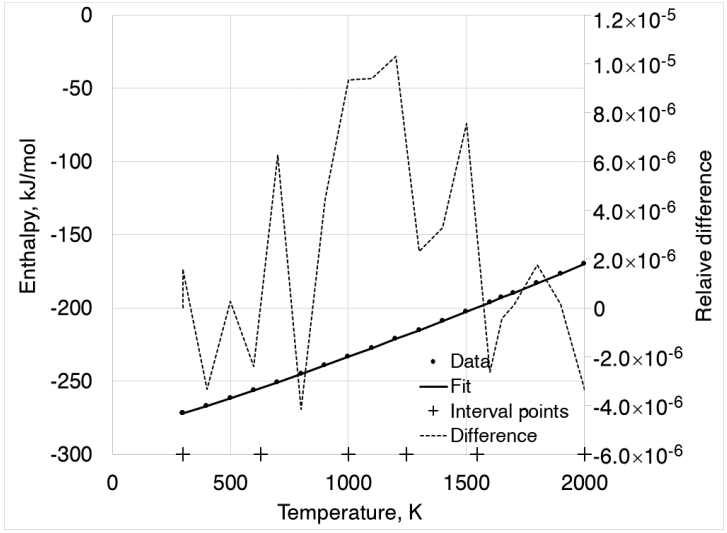


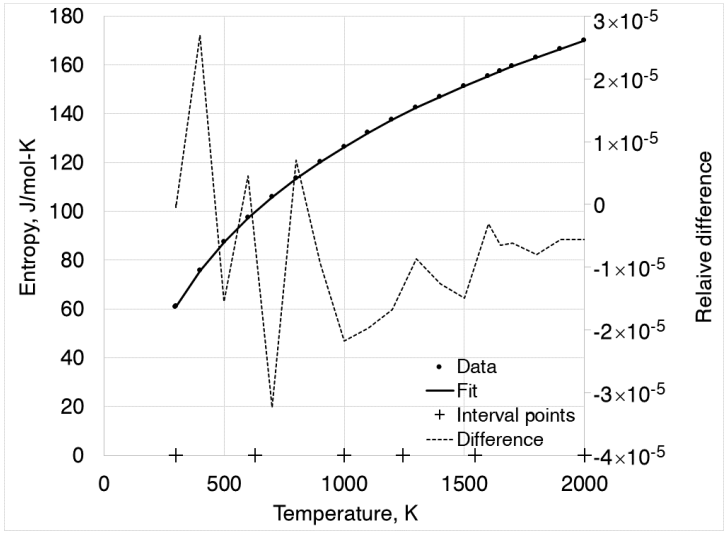


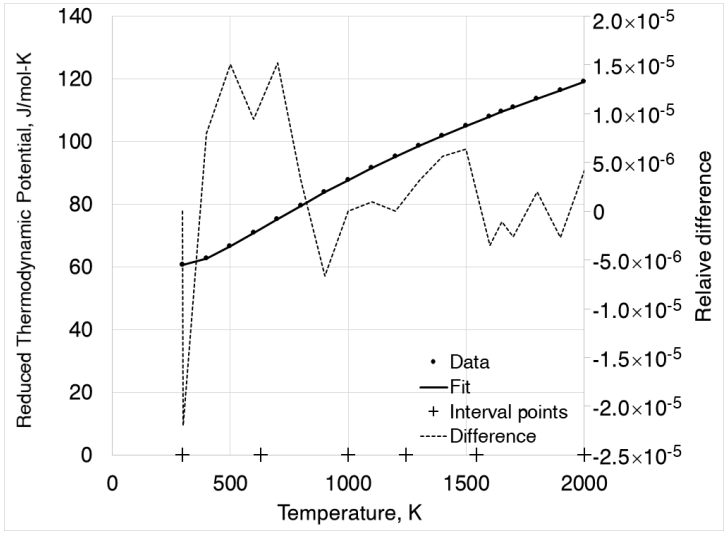
Iron Oxide (FeO) – Fe1O1(cr)

MW = 71.846 g/mol – $\Delta_f H^o$(298.15 K) = -272.044 kJ/mol – S^o(298.15 K) = 60.752 J/mol-K – p^o = 0.1 MPa



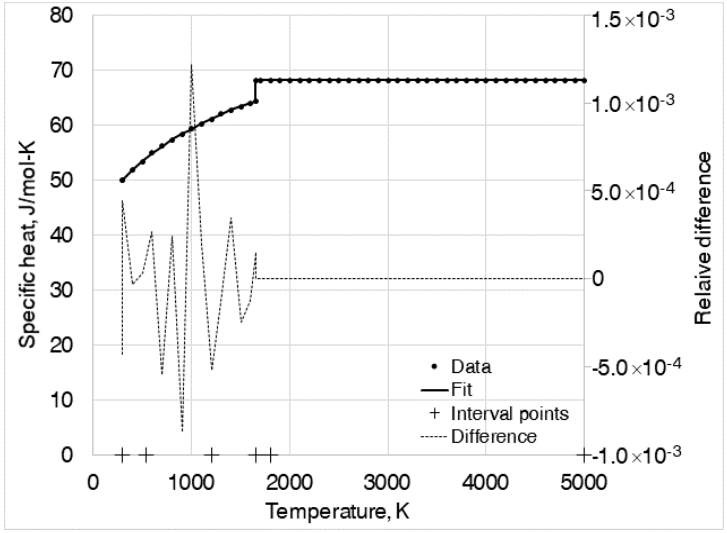


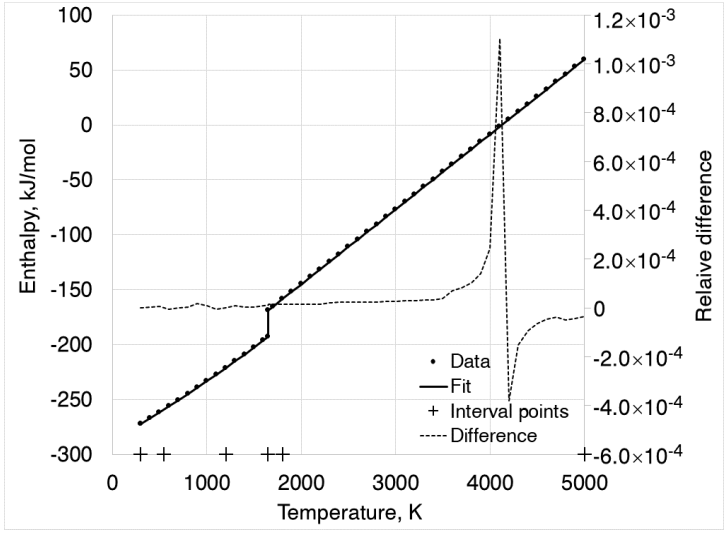


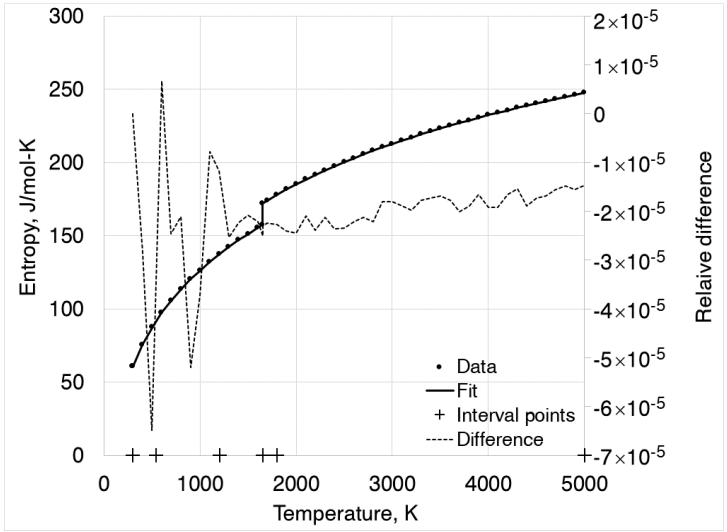


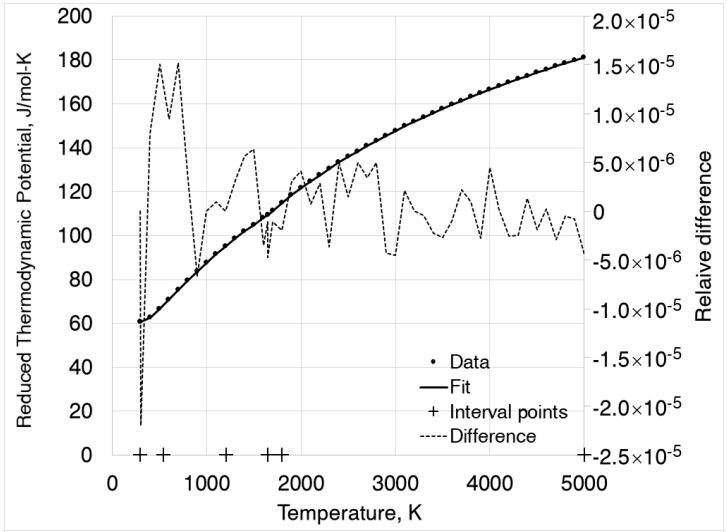
Iron Oxide (FeO) – Fe1O1(cr,l)

MW = 71.846 g/mol – $\Delta_f H^o$(298.15 K) = -272.044 kJ/mol – S^o(298.15 K) = 60.752 J/mol-K – p^o = 0.1 MPa



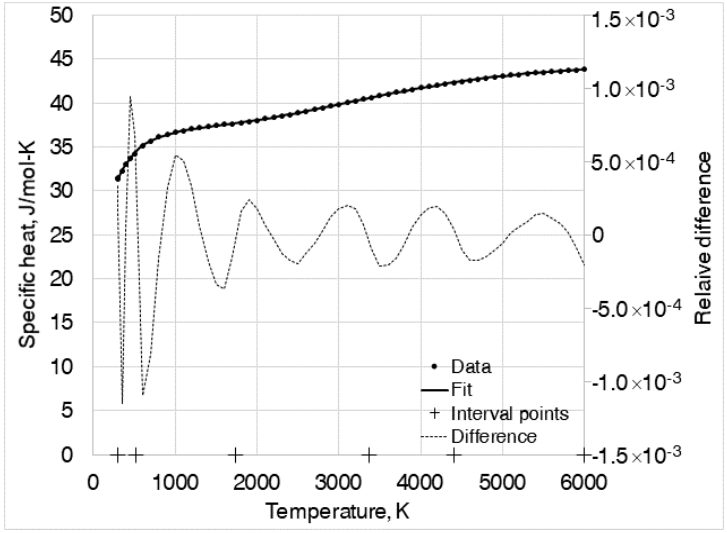


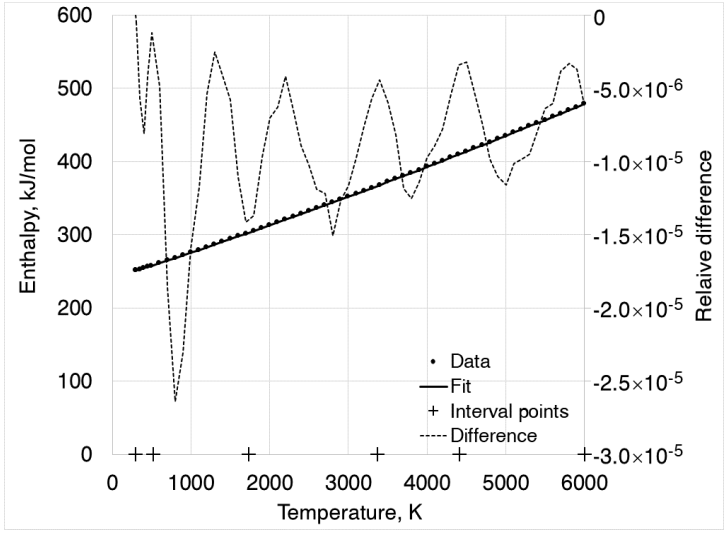


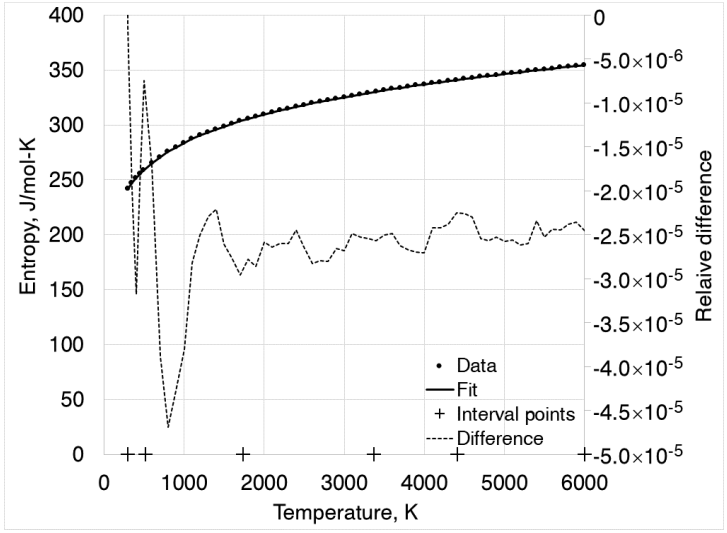


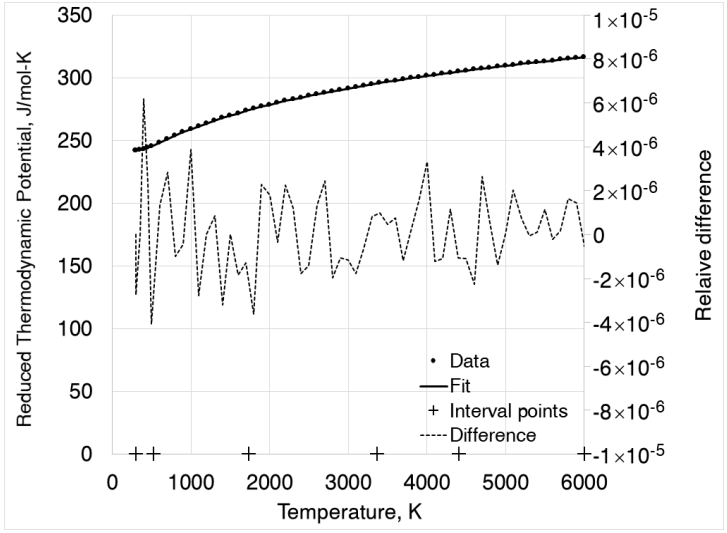
Iron Oxide (FeO) – Fe1O1(g)

MW = 71.846 g/mol – $\Delta_f H^o$(298.15 K) = 251.04 kJ/mol – S^o(298.15 K) = 241.924 J/mol-K – p^o = 0.1 MPa



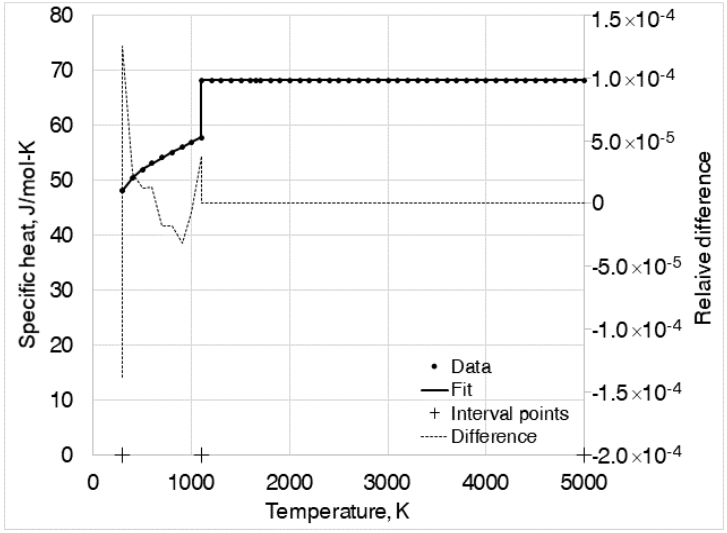


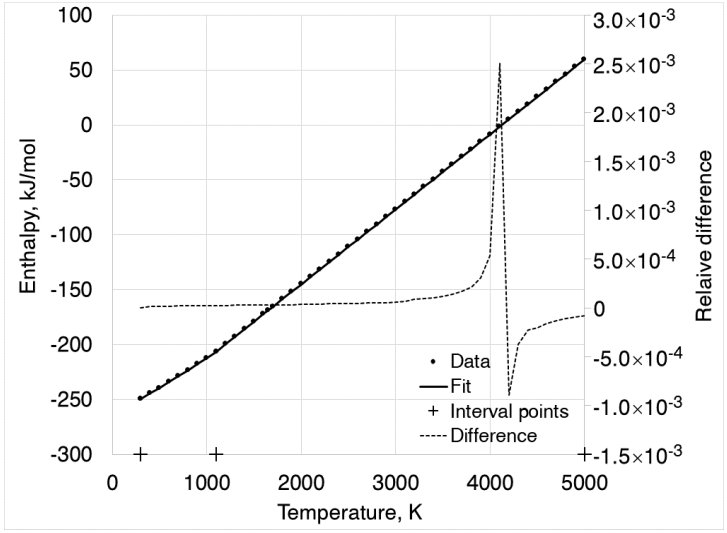


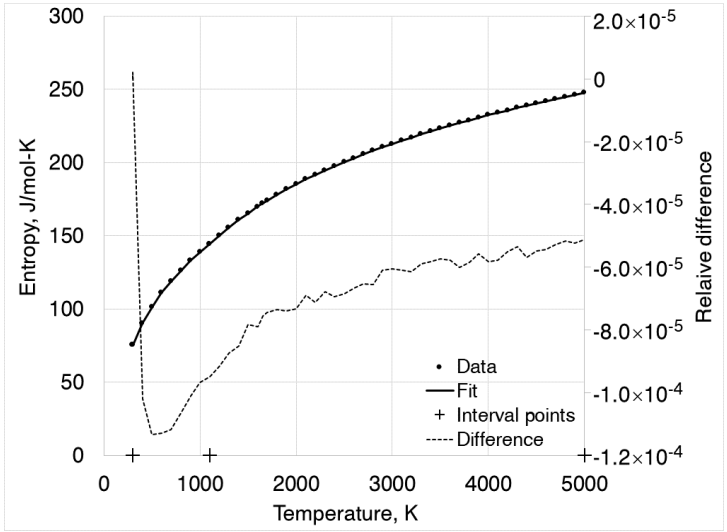


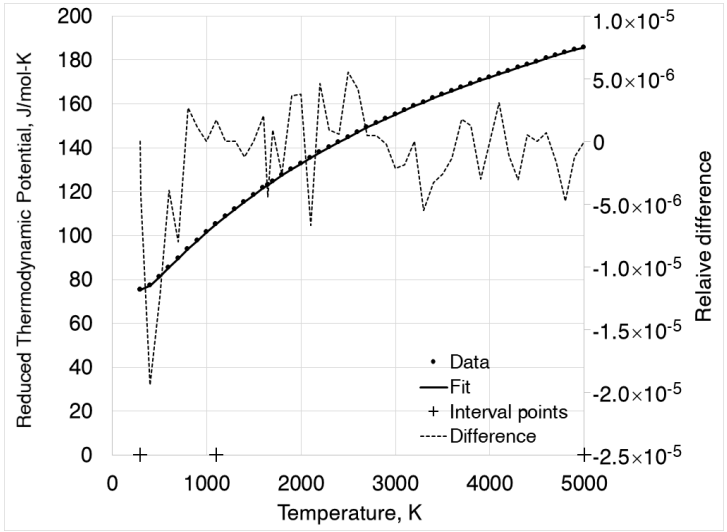
Iron Oxide (FeO) – Fe1O1(l)

MW = 71.846 g/mol – $\Delta_f H^o$(298.15 K) = -249.532 kJ/mol – S^o(298.15 K) = 75.415 J/mol-K – p^o = 0.1 MPa



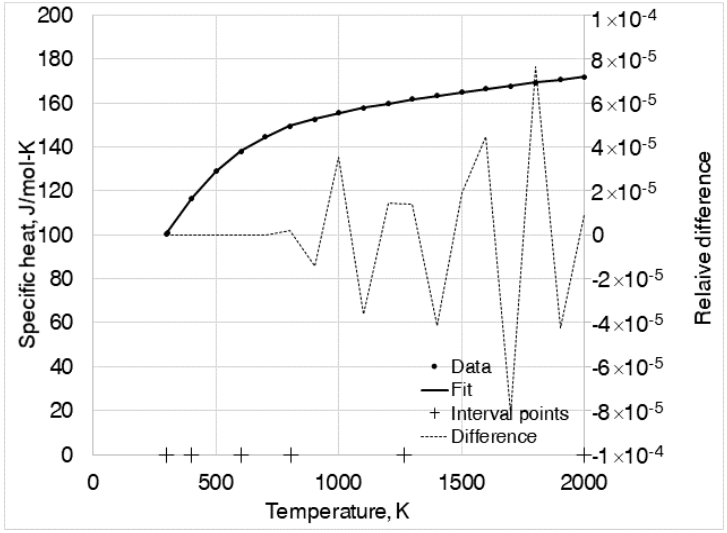


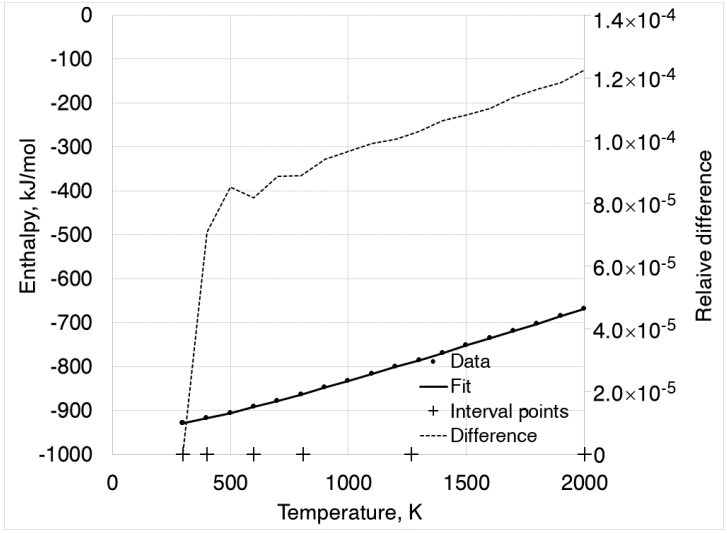


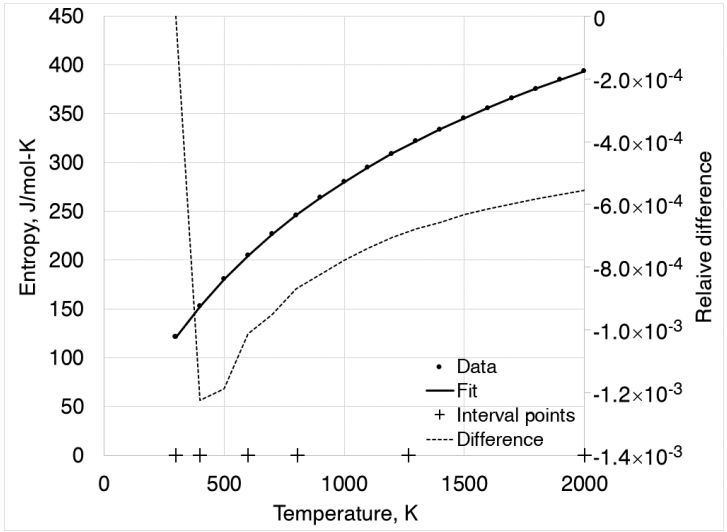


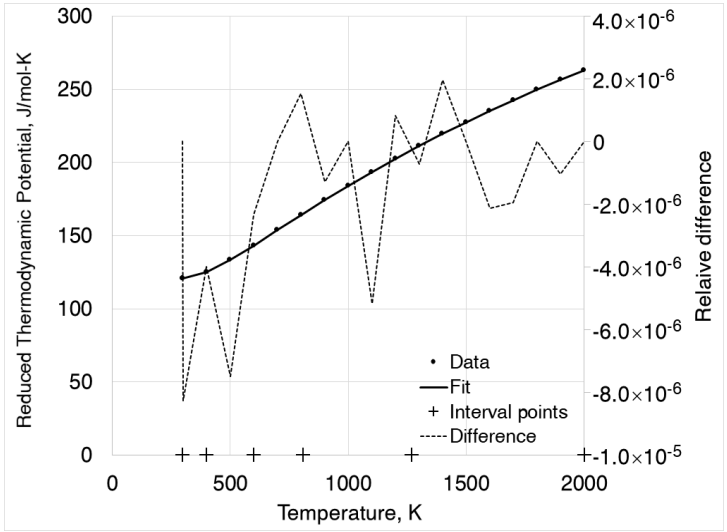
Iron Sulfate ($FeSO_4$) – Fe1O4S1(cr)

MW = 151.9 g/mol – $\Delta_f H^o$(298.15 K) = -928.848 kJ/mol – S^o(298.15 K) = 120.957 J/mol-K – p^o = 0.1 MPa



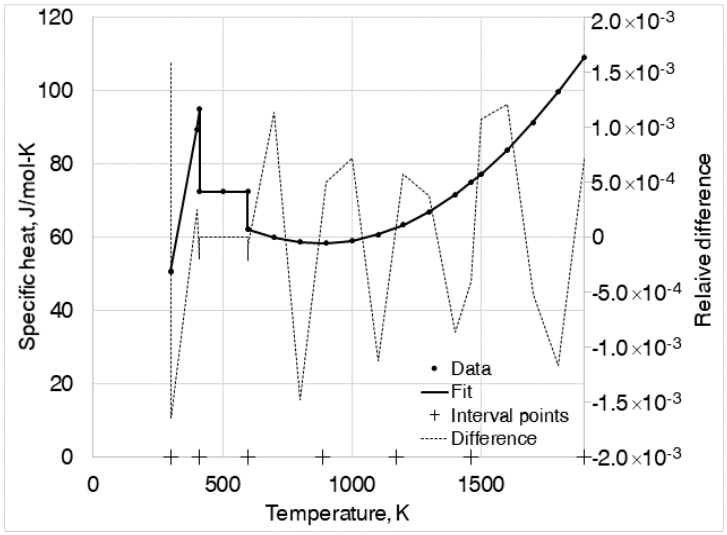


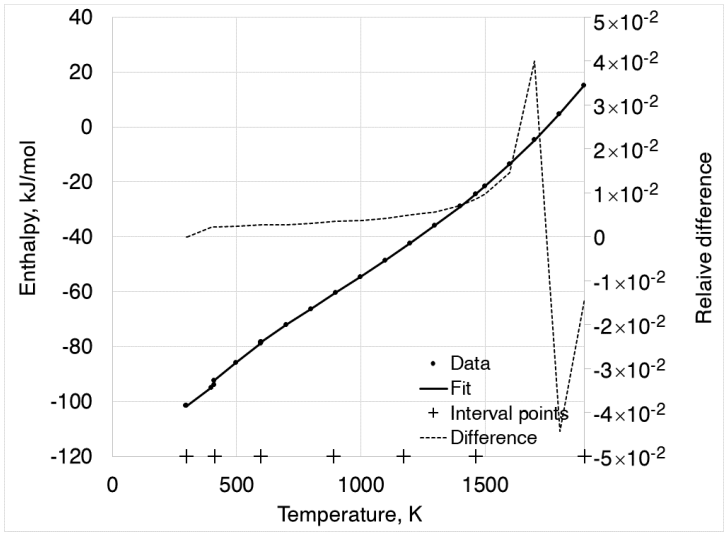


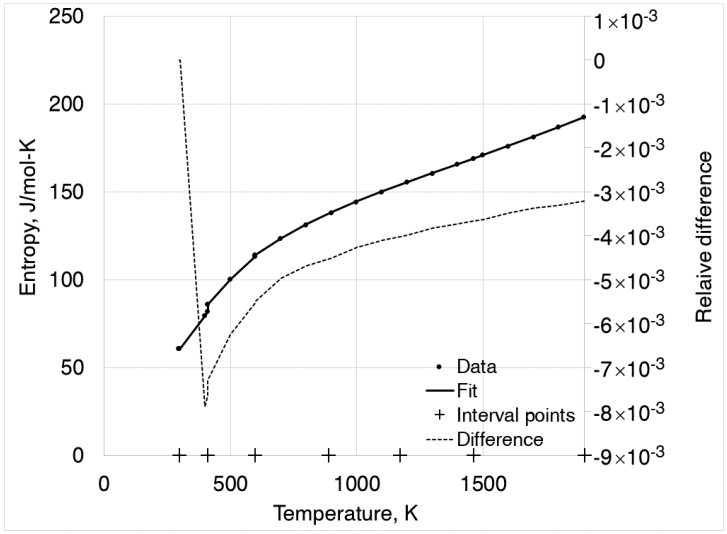


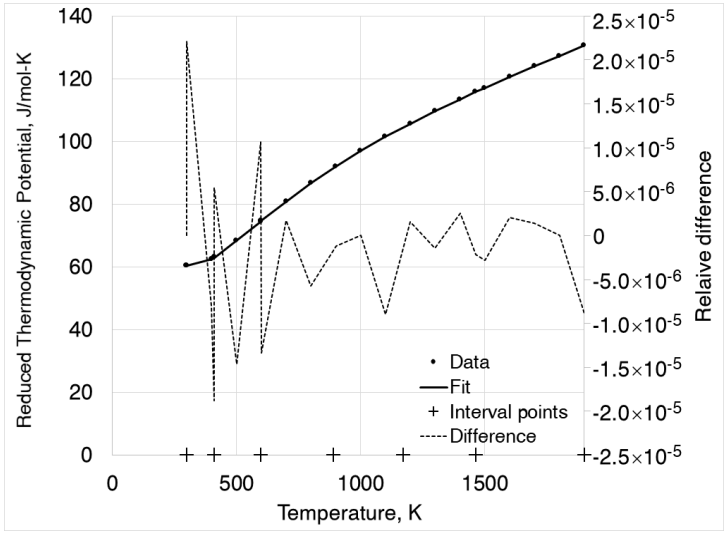
Iron Sulfide, Troilite (FeS) – Fe1S1(cr)

MW = 87.91 g/mol – $\Delta_f H^o$(298.15 K) = -101.671 kJ/mol – S^o(298.15 K) = 60.321 J/mol-K – p^o = 0.1 MPa



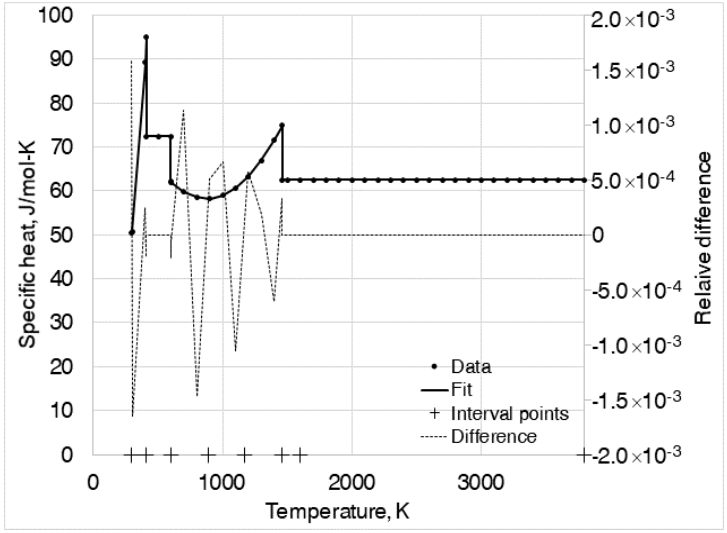


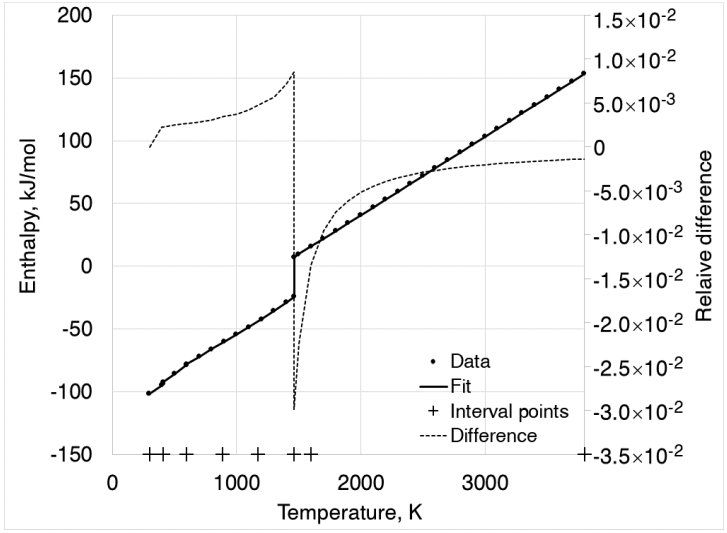


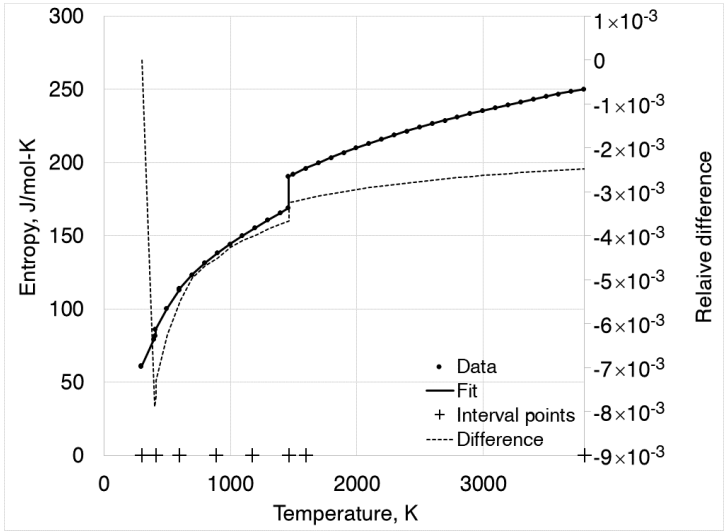


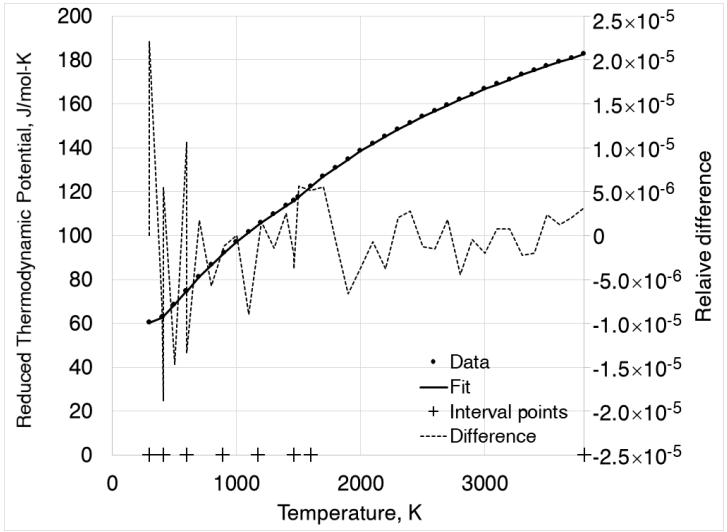
Iron Sulfide, Troilite (FeS) – Fe1S1(cr,l)

MW = 87.91 g/mol – $\Delta_f H^o$(298.15 K) = -101.671 kJ/mol – S^o(298.15 K) = 60.321 J/mol-K – p^o = 0.1 MPa



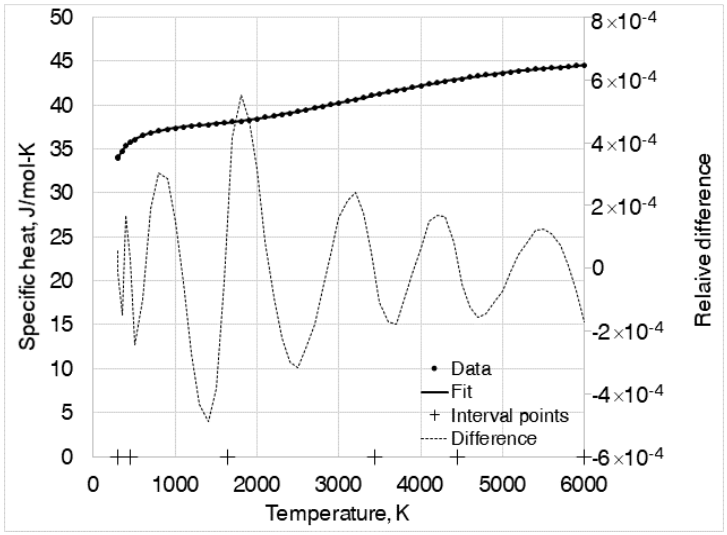


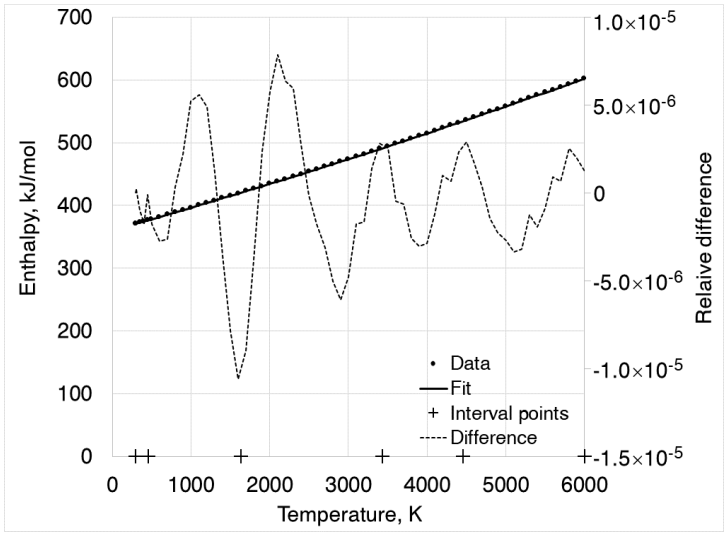


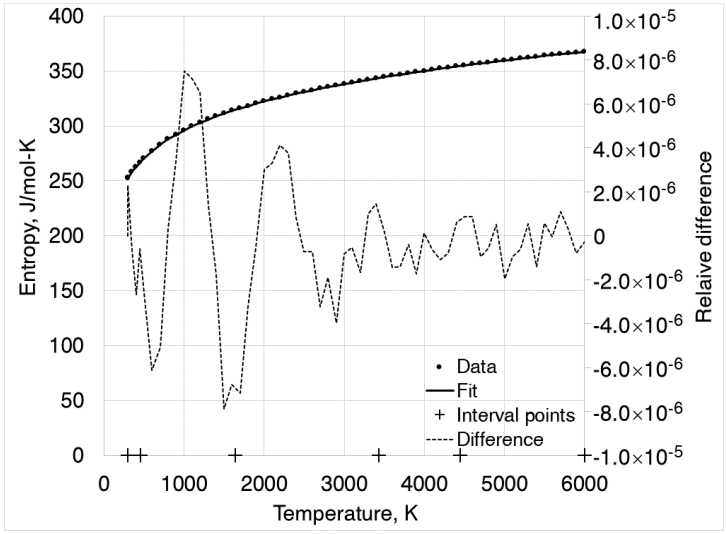


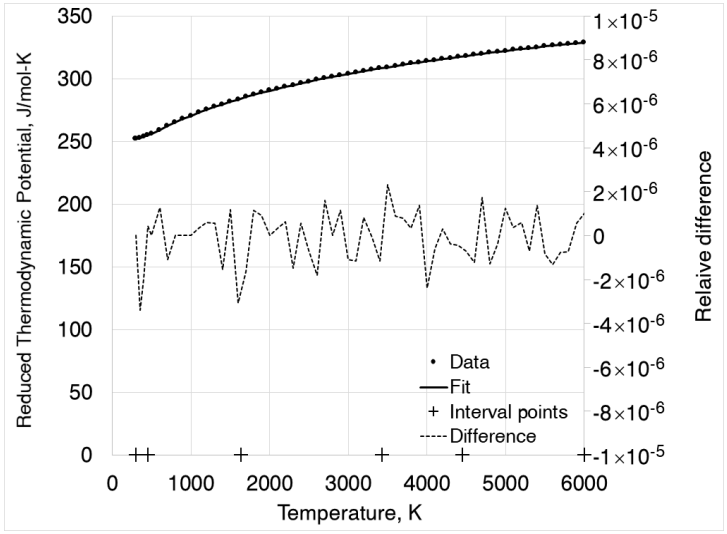
Iron Sulfide (FeS) – Fe1S1(g)

MW = 87.91 g/mol – $\Delta_f H^o$(298.15 K) = 370.767 kJ/mol – S^o(298.15 K) = 252.344 J/mol-K – p^o = 0.1 MPa



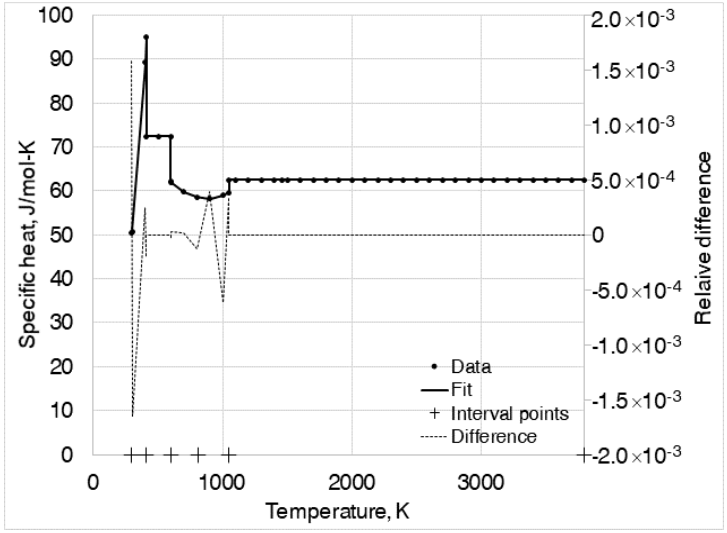


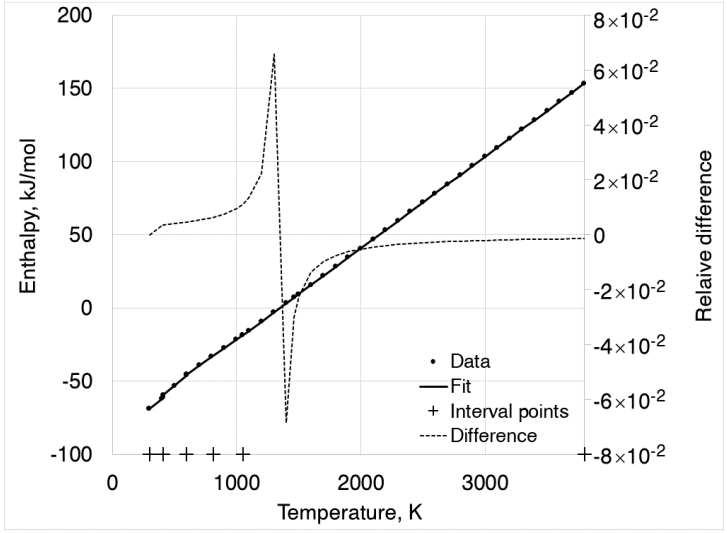


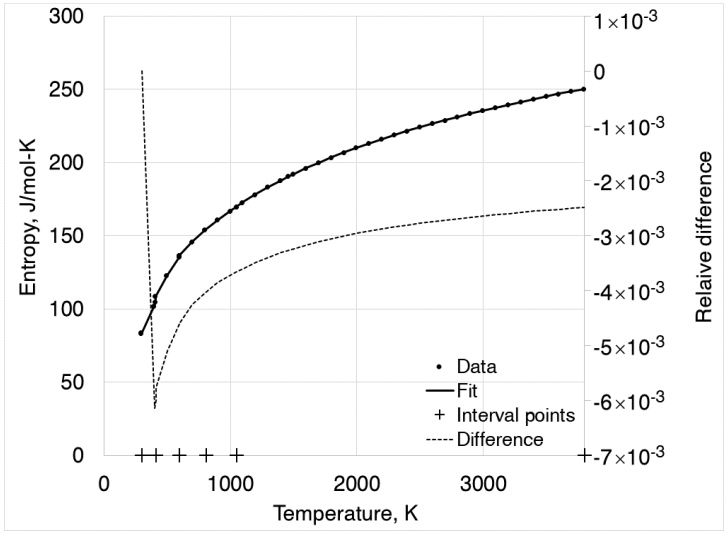


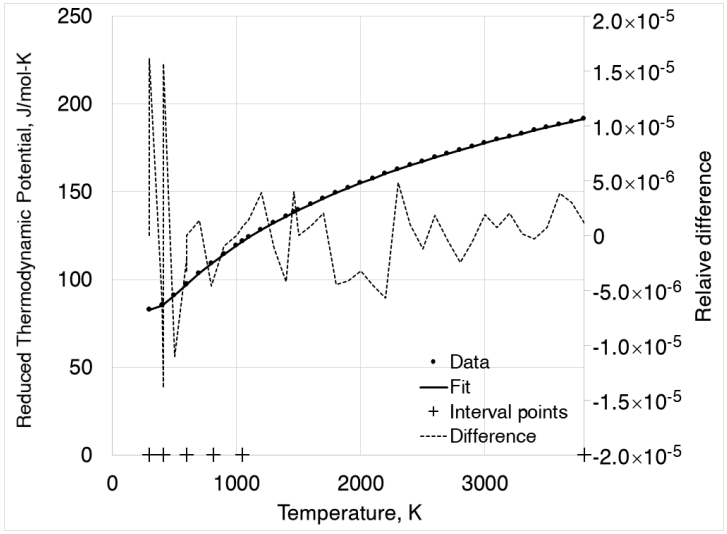
Iron Sulfide (FeS) – Fe1S1(l)

MW = 87.91 g/mol – $\Delta_f H^o$(298.15 K) = -68.811 kJ/mol – S^o(298.15 K) = 82.811 J/mol-K – p^o = 0.1 MPa



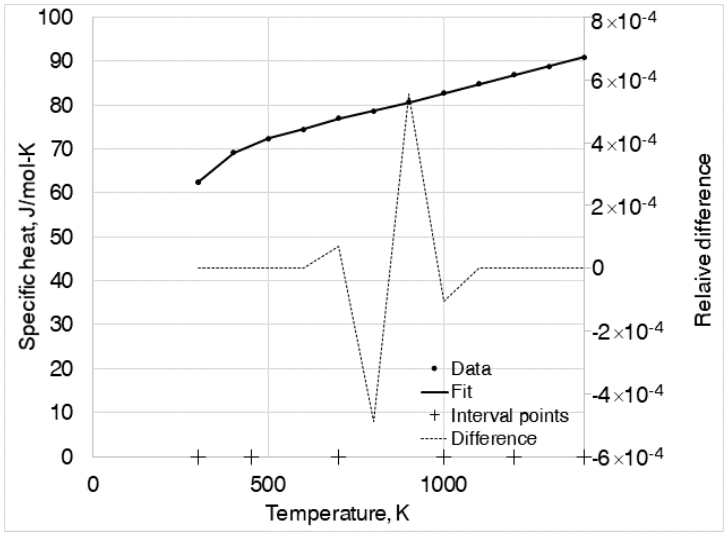


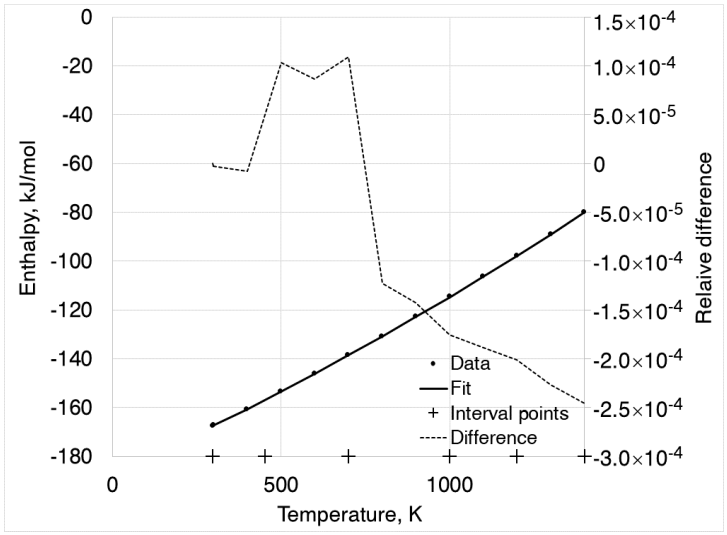


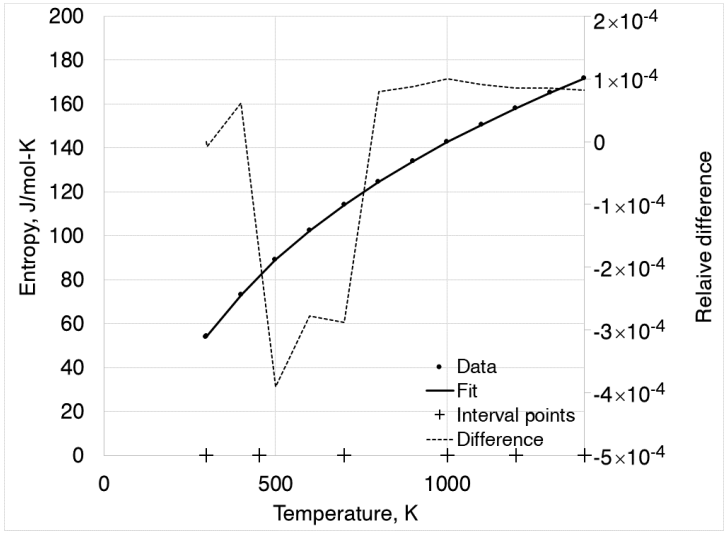


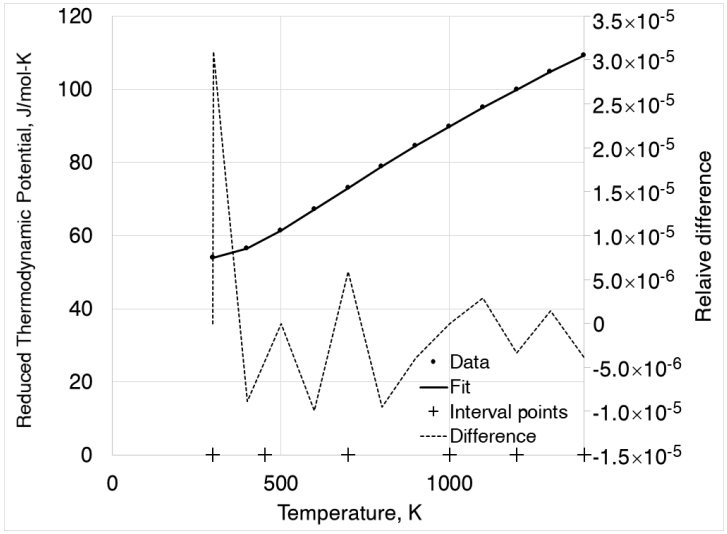
Iron Sulfide, Marcasite (FeS_2) – Fe1S2(cr)_01

MW = 119.97 g/mol – $\Delta_f H^o$(298.15 K) = -167.36 kJ/mol – S^o(298.15 K) = 53.865 J/mol-K – p^o = 0.1 MPa



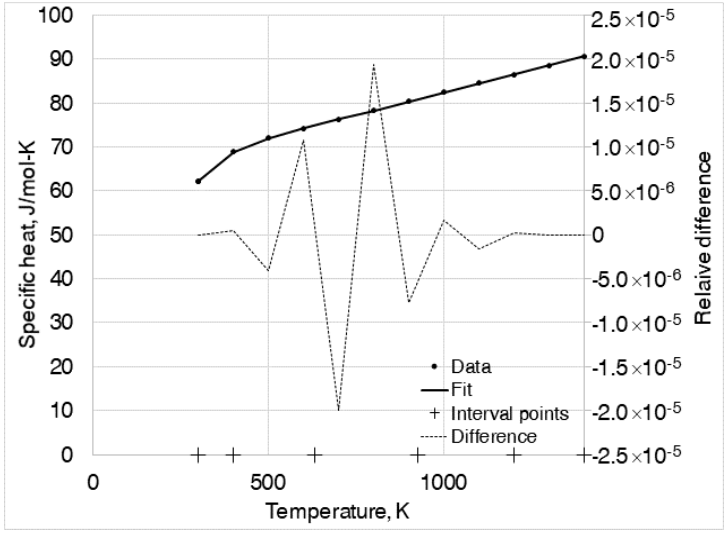


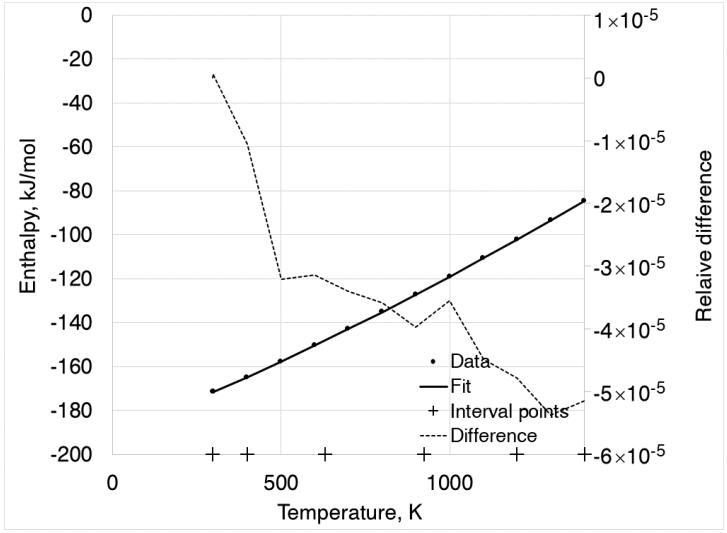


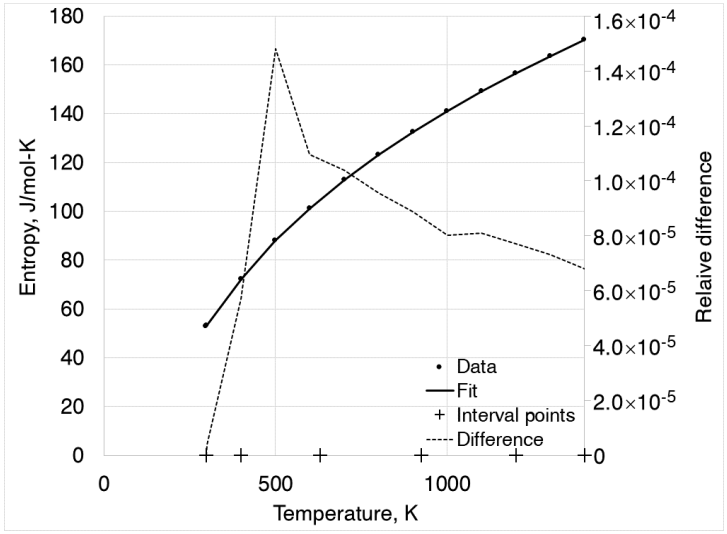


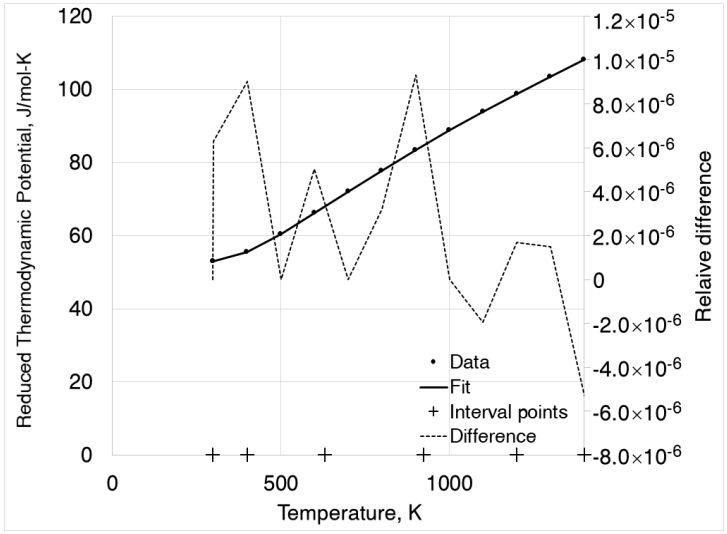
Iron Sulfide, Pyrite (FeS_2) – Fe1S2(cr)_02

MW = 119.97 g/mol – $\Delta_f H^o$(298.15 K) = -171.544 kJ/mol – S^o(298.15 K) = 52.916 J/mol-K – p^o = 0.1 MPa



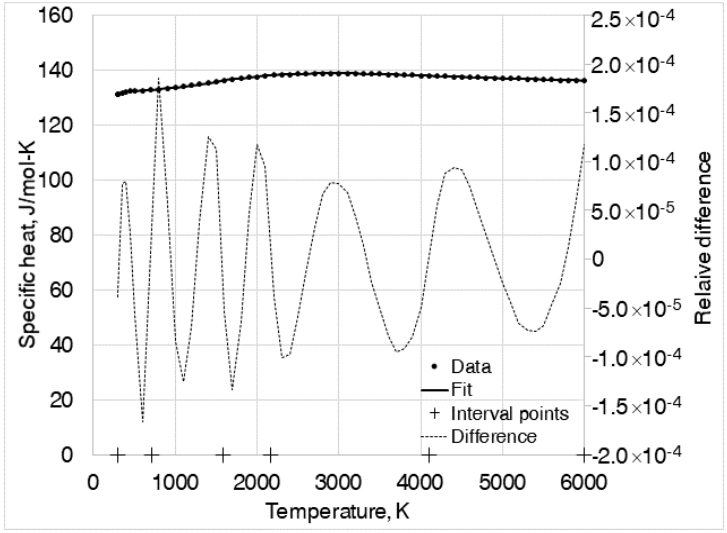


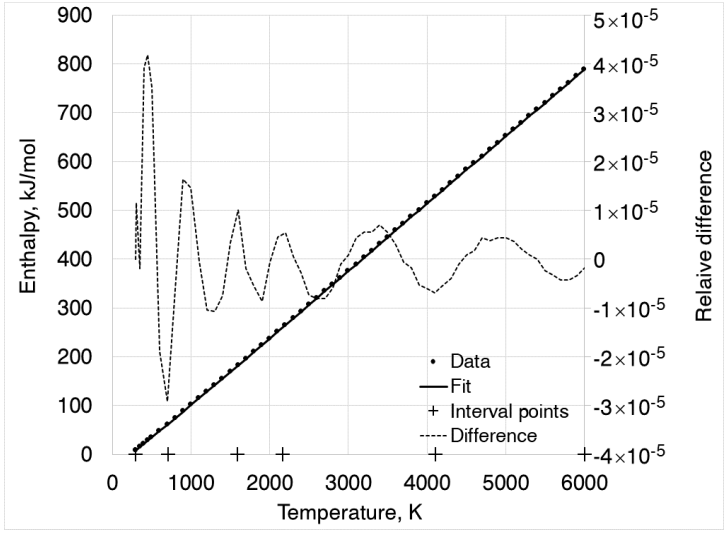


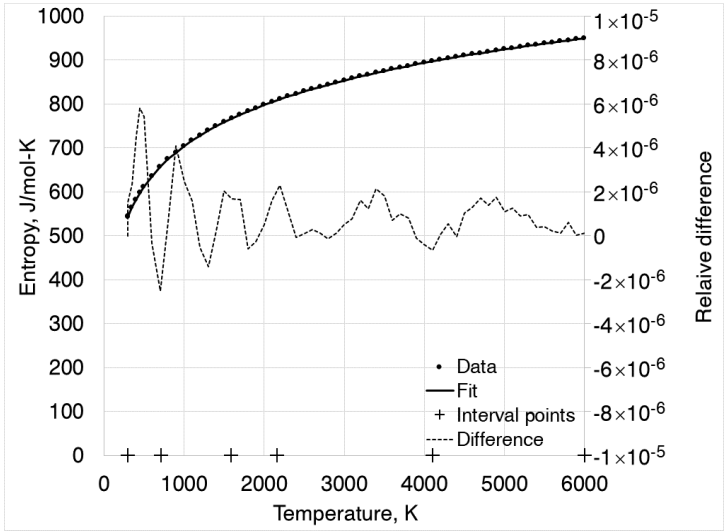


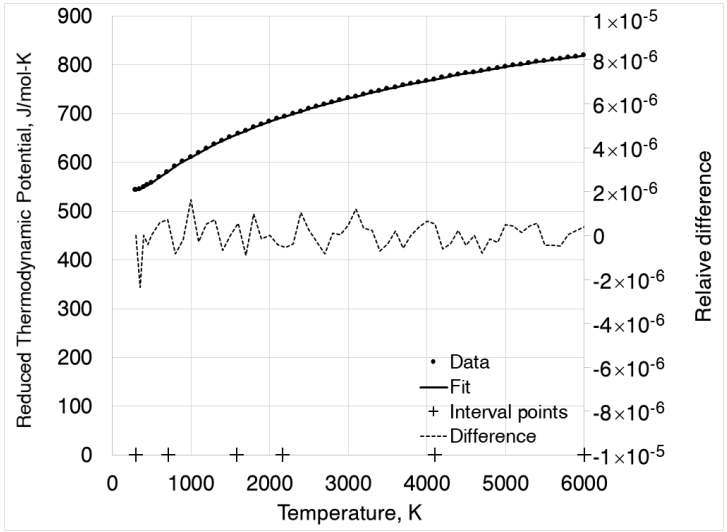
Iron Iodide ($(FeI_2)_2$) – Fe2I4(g)

MW = 619.312 g/mol – $\Delta_f H^o$(298.15 K) = 8.368 kJ/mol – S^o(298.15 K) = 543.471 J/mol-K – p^o = 0.1 MPa



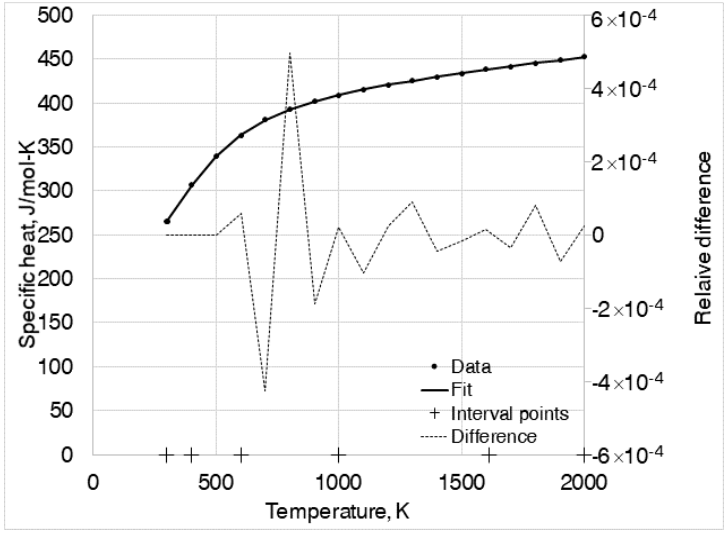


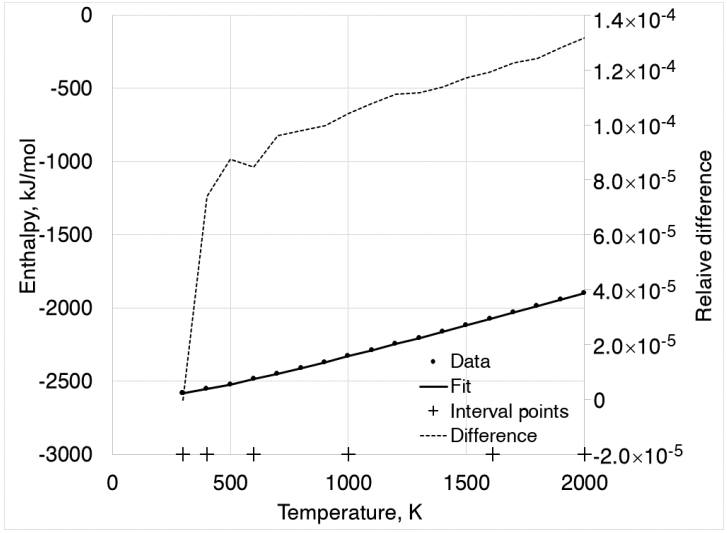


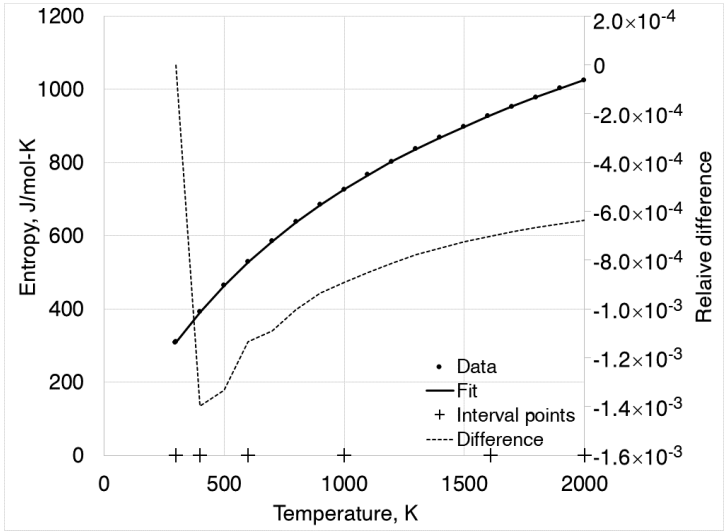


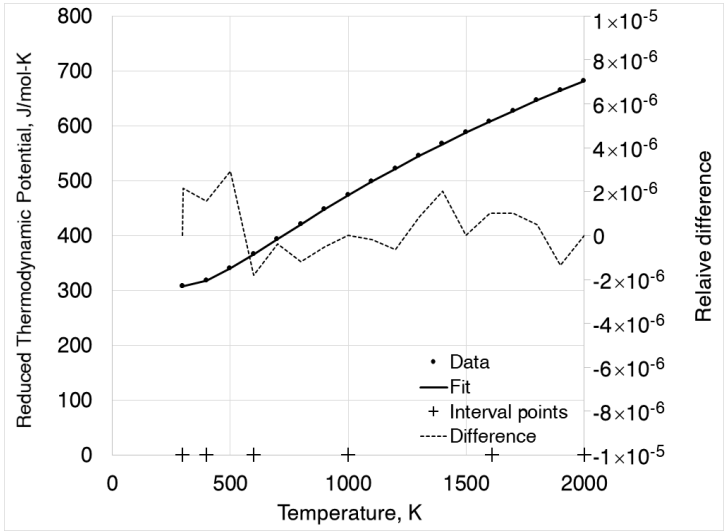
Iron Sulfate ($Fe_2(SO_4)_3$) – Fe2O12S3(cr)

MW = 399.87 g/mol – $\Delta_f H^o$(298.15 K) = -2582.992 kJ/mol – S^o(298.15 K) = 307.524 J/mol-K – p^o = 0.1 MPa



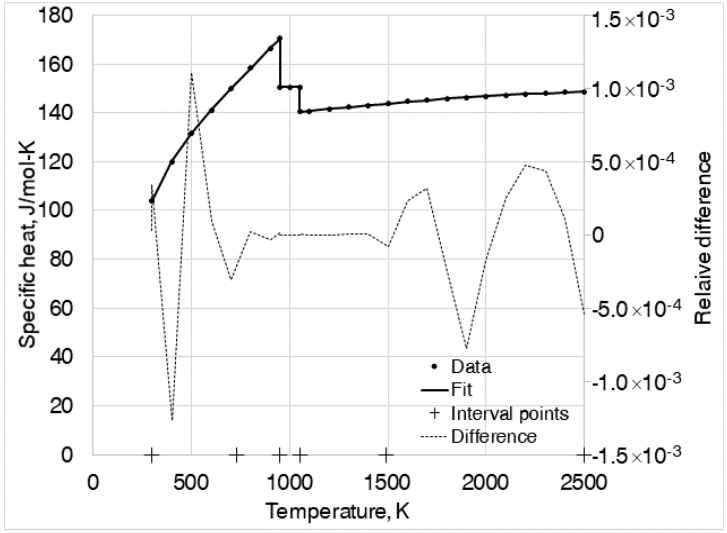


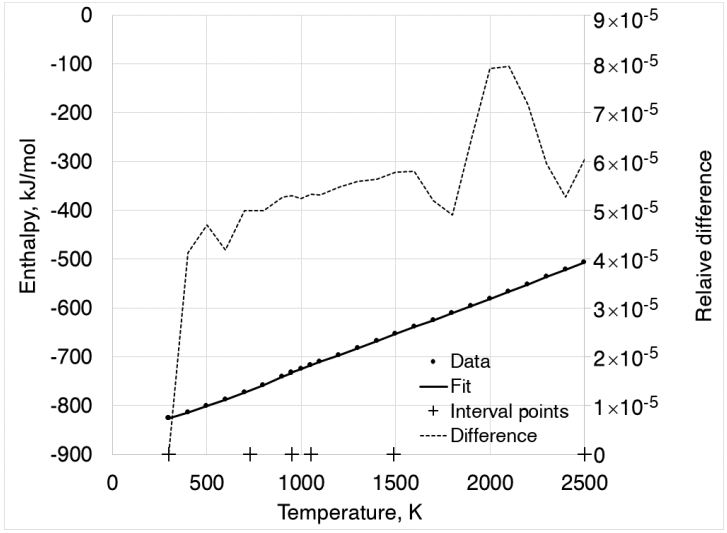


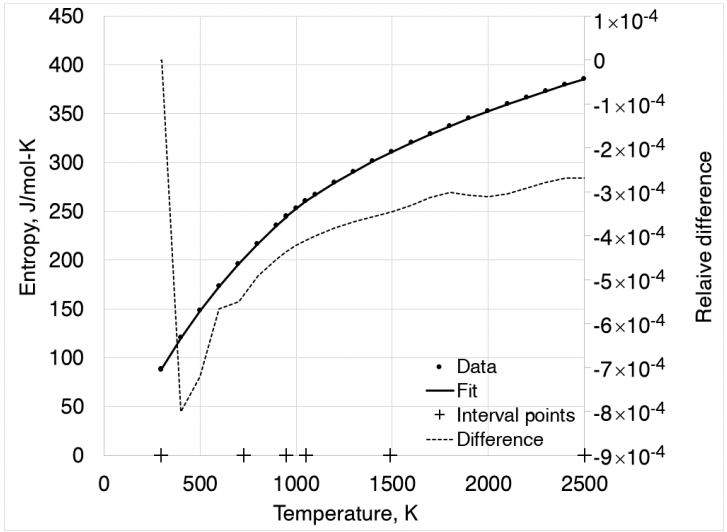


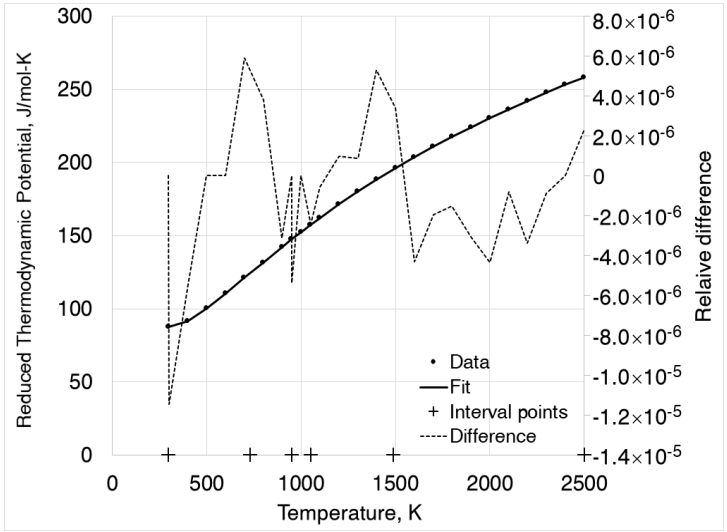
Iron Oxide, Hematite (Fe_2O_3) – Fe2O3(cr)

MW = 159.692 g/mol – $\Delta_f H^o$(298.15 K) = -825.503 kJ/mol – S^o(298.15 K) = 87.4 J/mol-K – p^o = 0.1 MPa



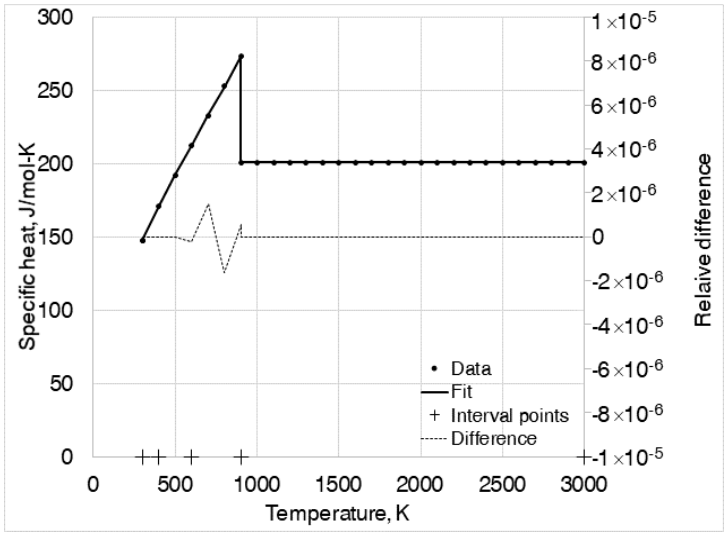


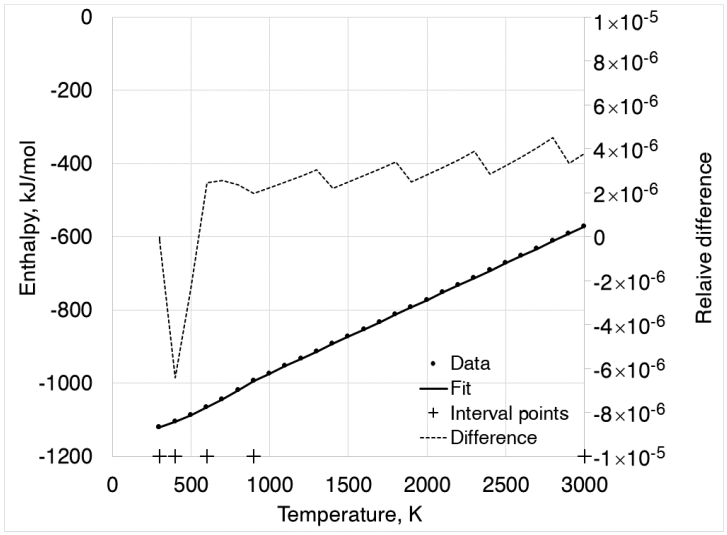


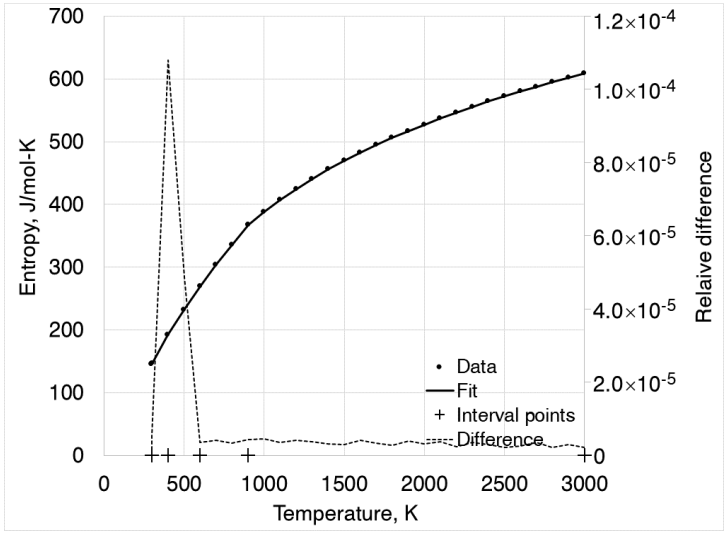


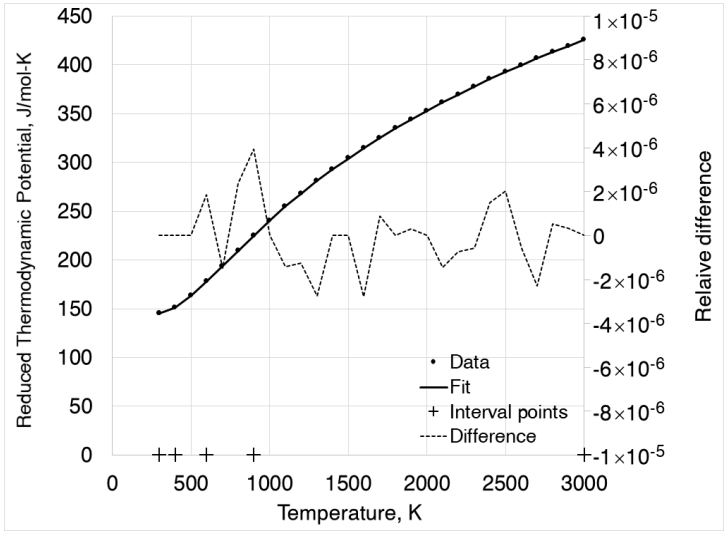
Iron Oxide, Magnetite (Fe_3O_4) – Fe3O4(cr)

MW = 231.539 g/mol – $\Delta_f H^o$(298.15 K) = -1120.894 kJ/mol – S^o(298.15 K) = 145.266 J/mol-K – p^o = 0.1 MPa



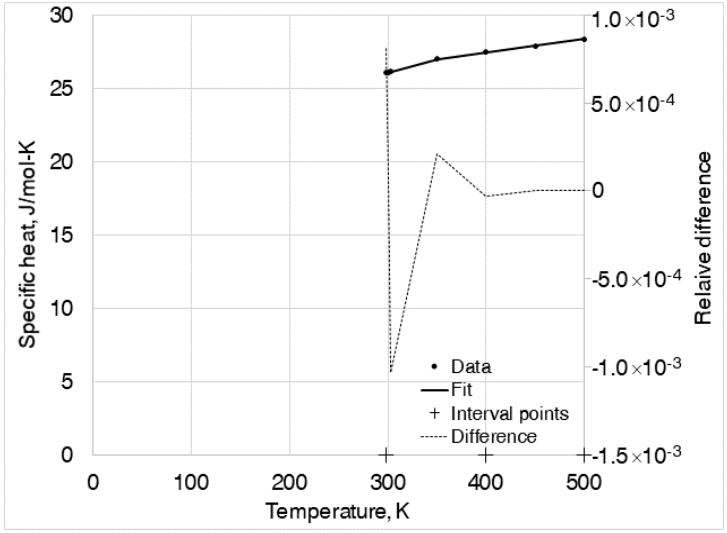


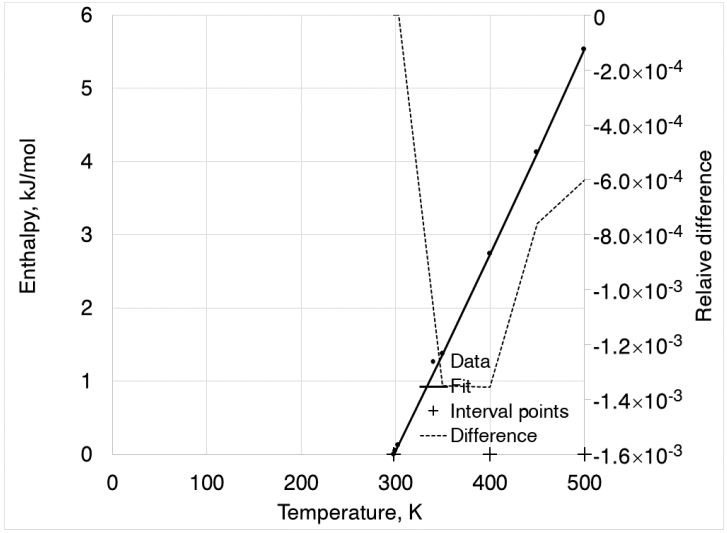


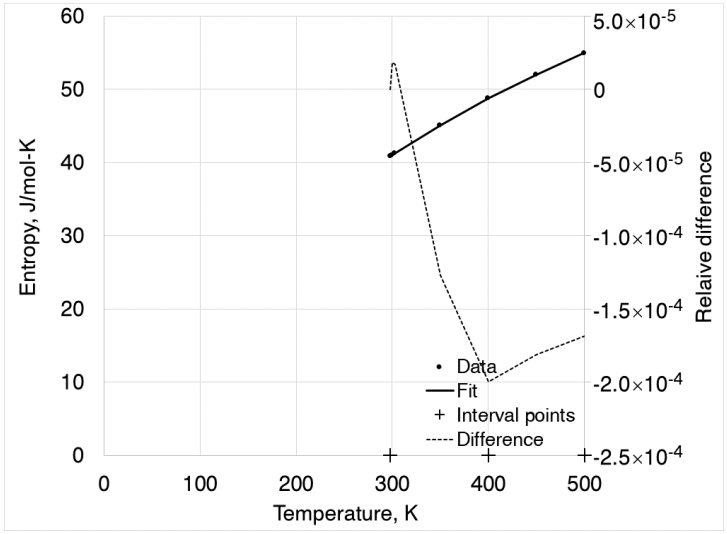


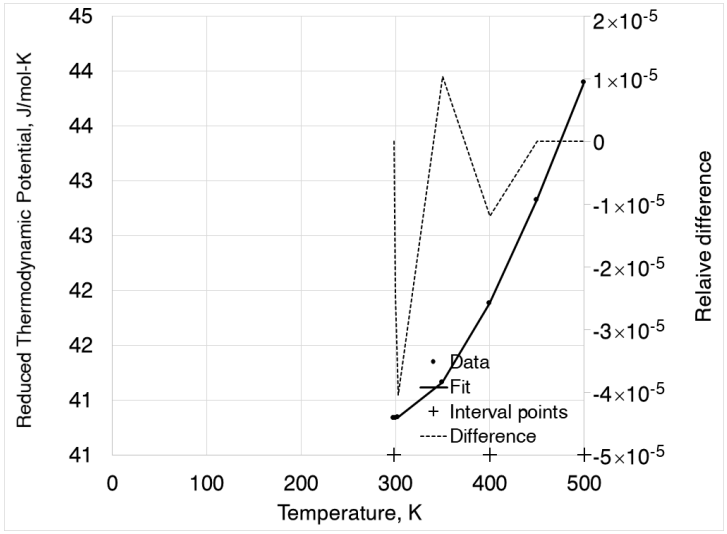
Gallium (Ga) – Ga1(cr)

MW = 69.72 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 40.838 J/mol-K – p^o = 0.1 MPa



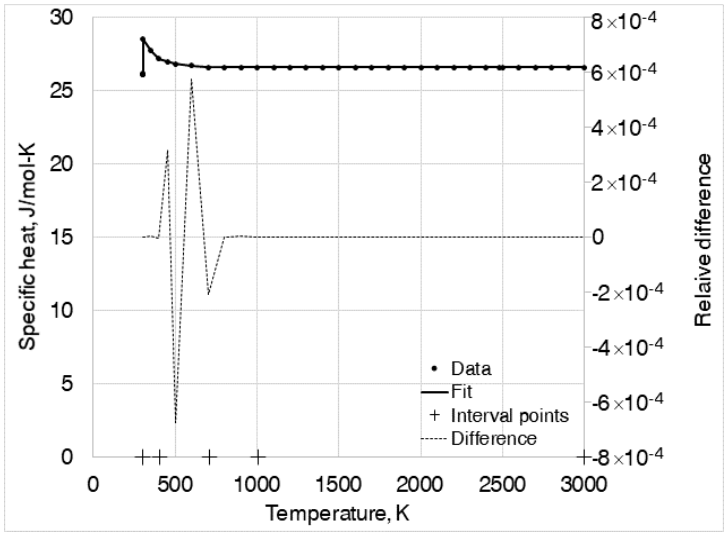


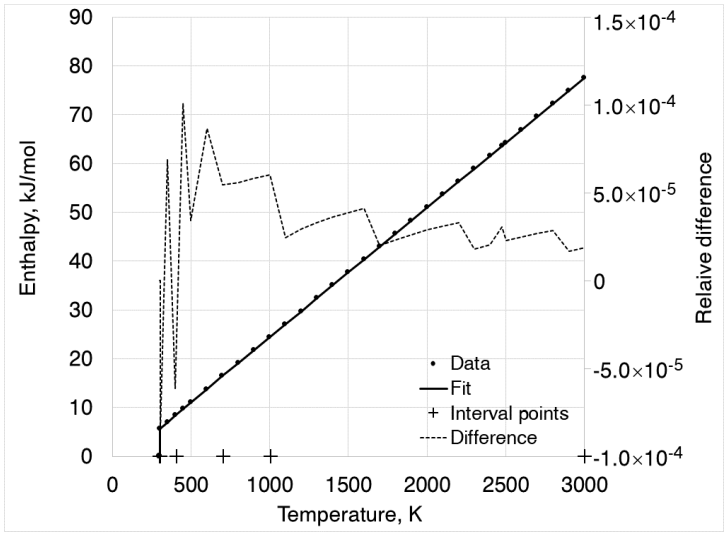


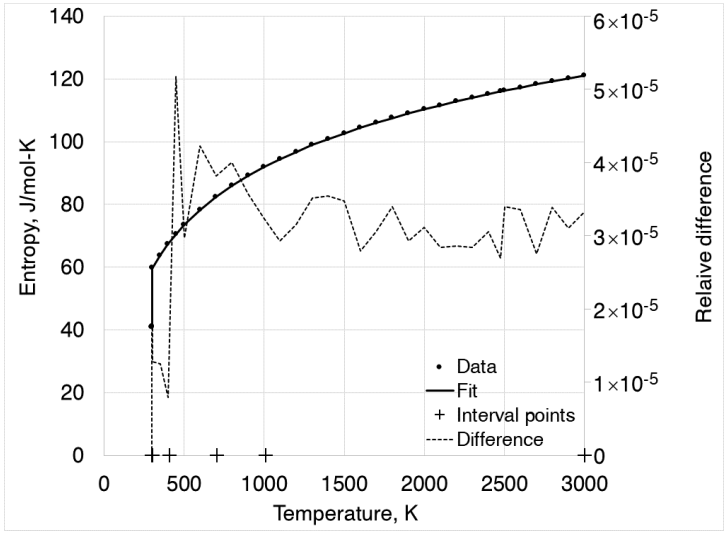


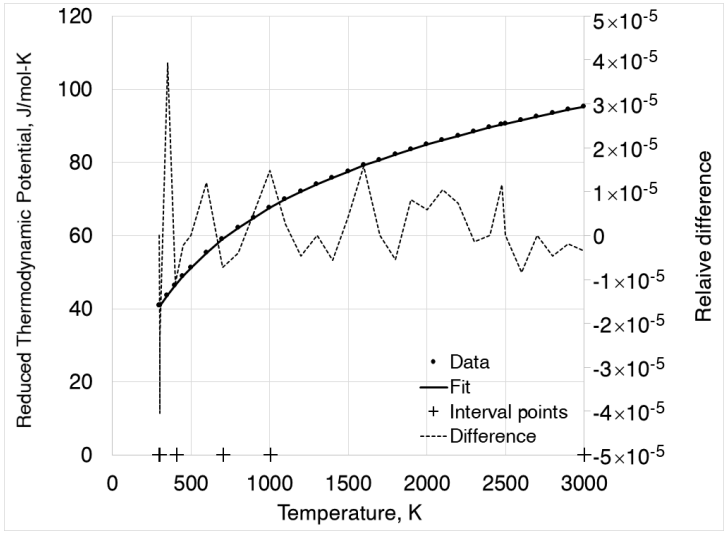
Gallium (Ga) – Ga1(cr,l)

MW = 69.72 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 40.838 J/mol-K – p^o = 0.1 MPa



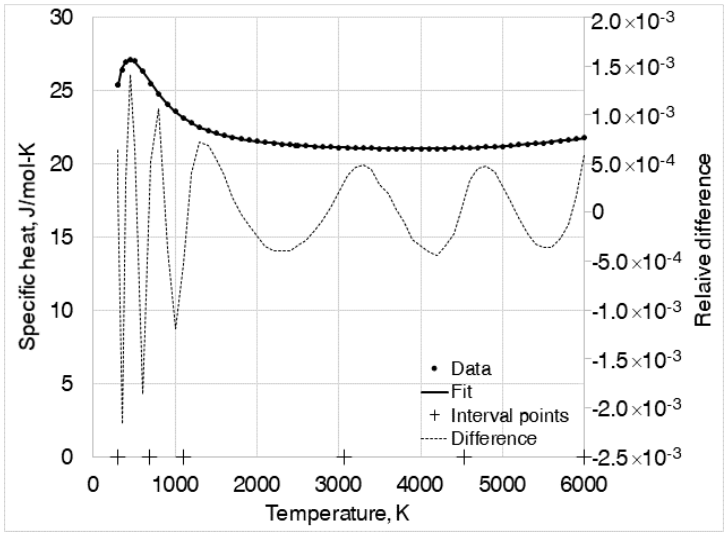


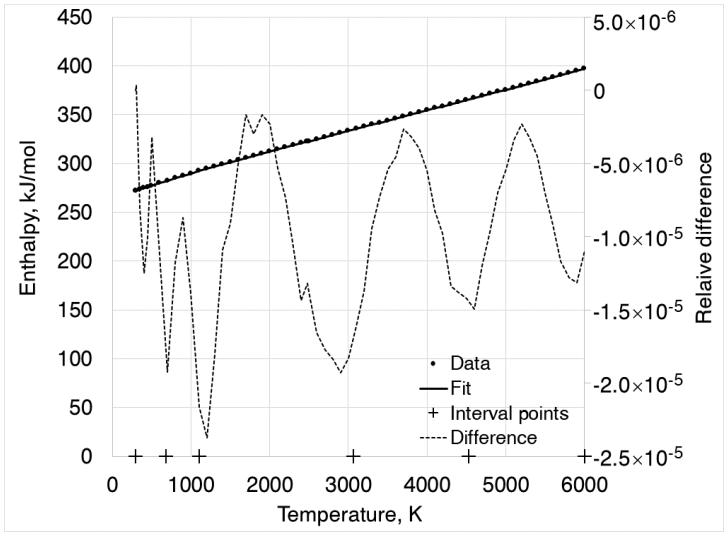


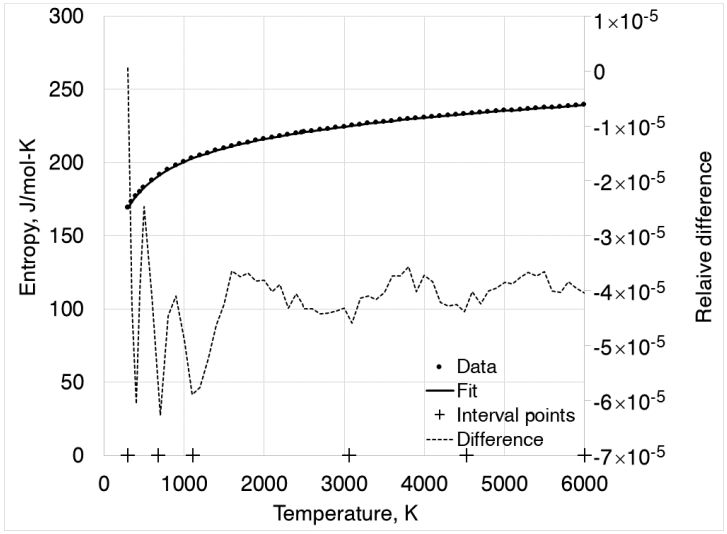


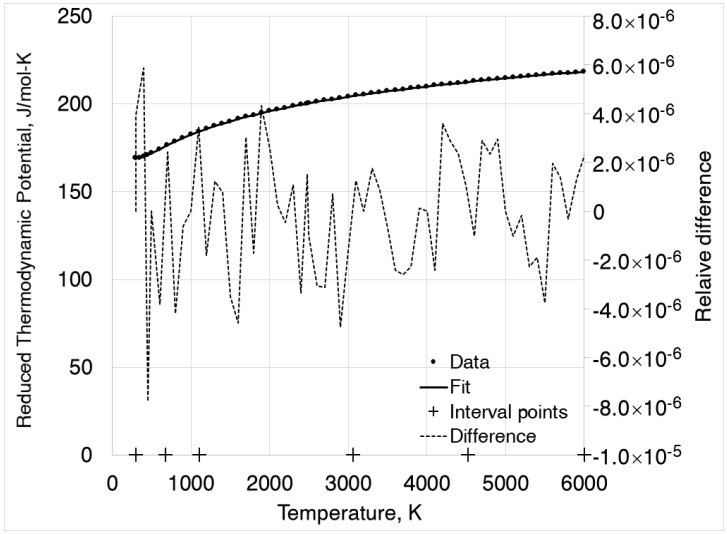
Gallium (Ga) – Ga1(g)

MW = 69.72 g/mol – $\Delta_f H^o$(298.15 K) = 271.96 kJ/mol – S^o(298.15 K) = 169.042 J/mol-K – p^o = 0.1 MPa



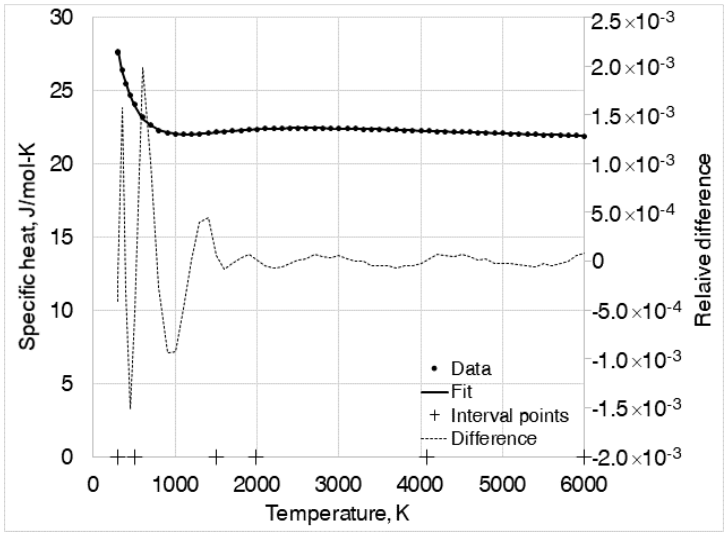


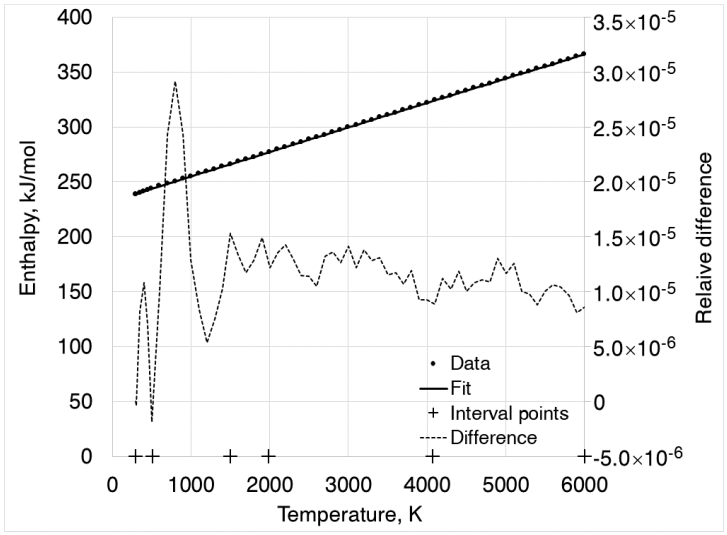


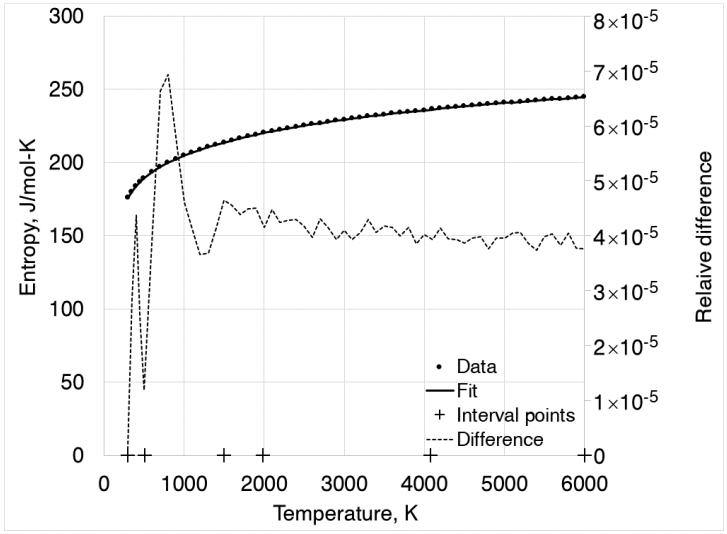


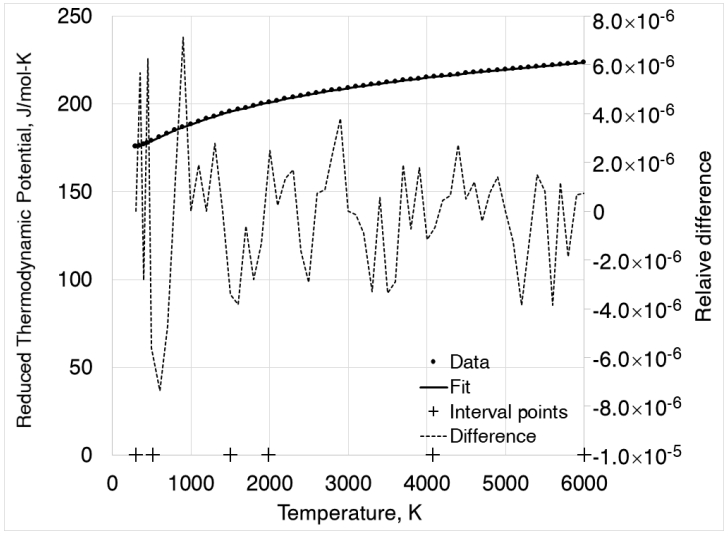
Gallium, Ion (Ga$^-$) – Ga1-(g)

MW = 69.72 g/mol – $\Delta_f H^o$(298.15 K) = 238.531 kJ/mol – S^o(298.15 K) = 175.801 J/mol-K – p^o = 0.1 MPa



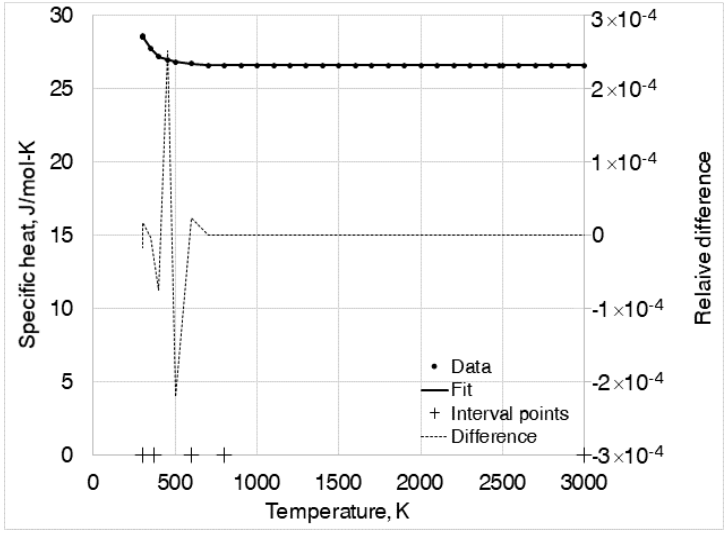


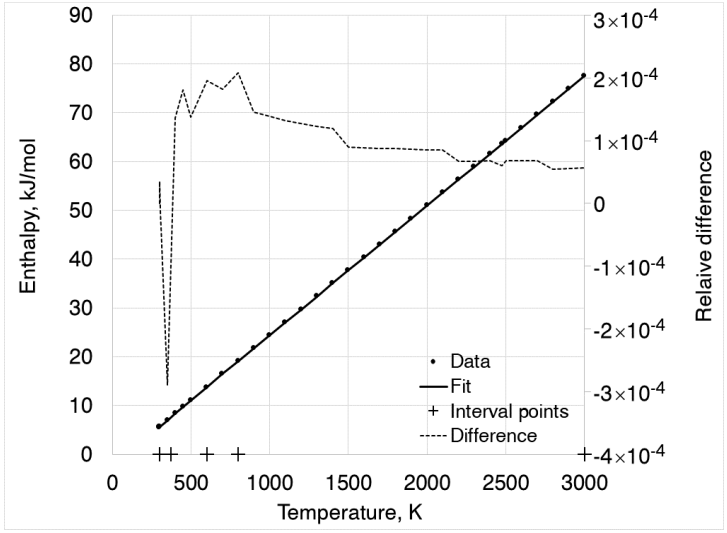


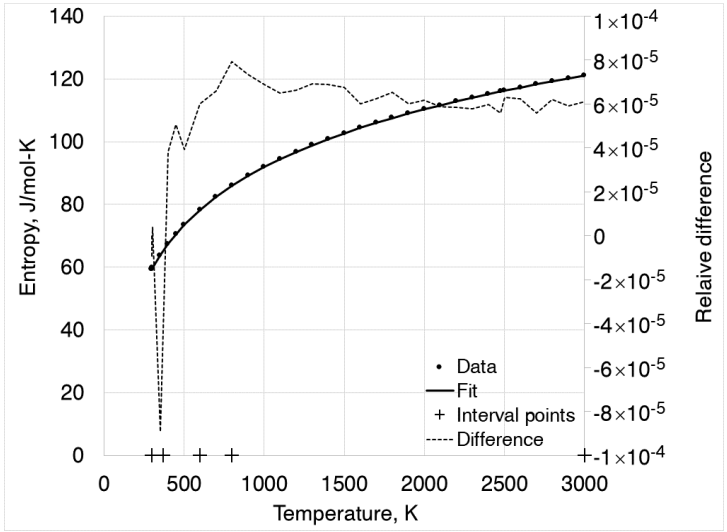


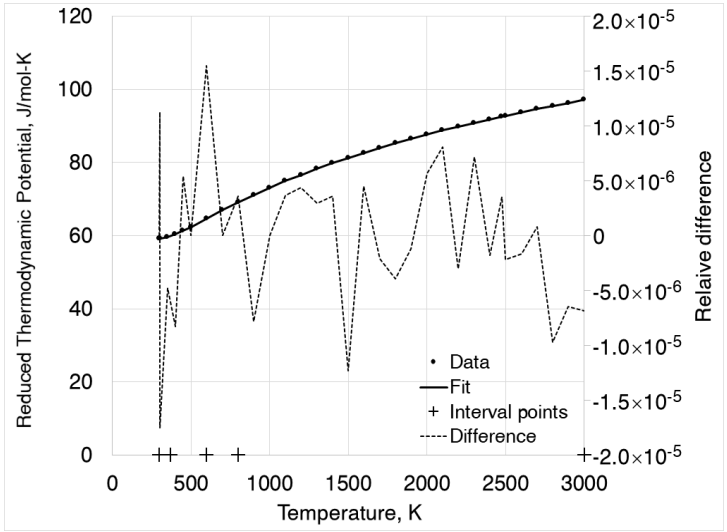
Gallium (Ga) – Ga1(l)

MW = 69.72 g/mol – $\Delta_f H^o$(298.15 K) = 5.578 kJ/mol – S^o(298.15 K) = 59.253 J/mol-K – p^o = 0.1 MPa



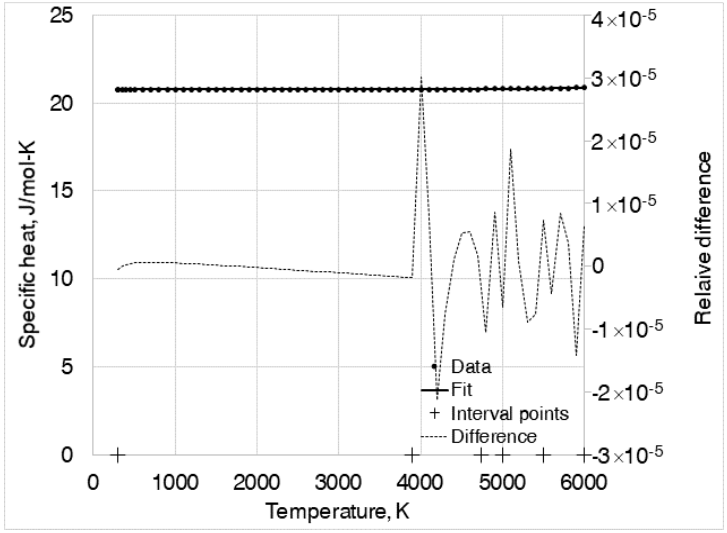


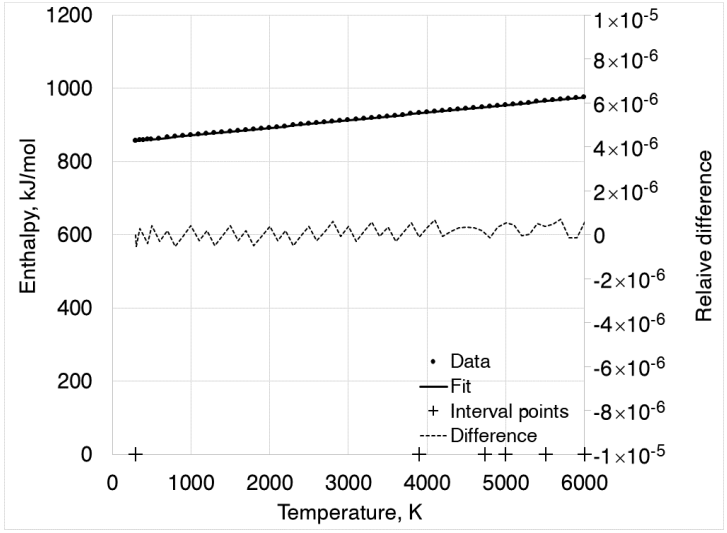


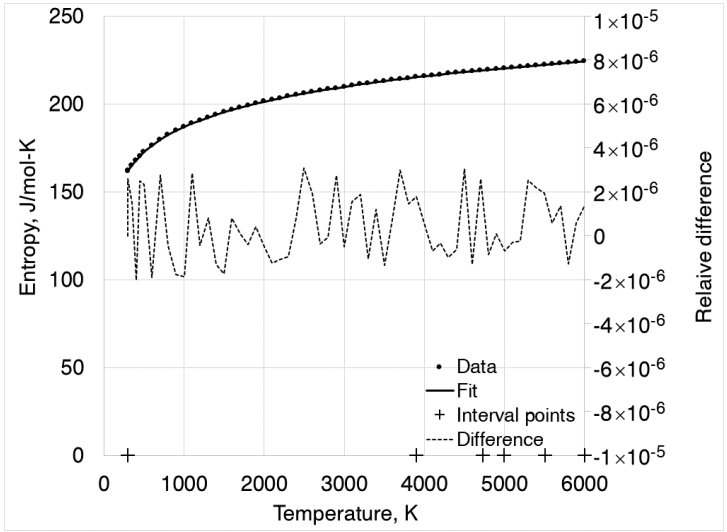


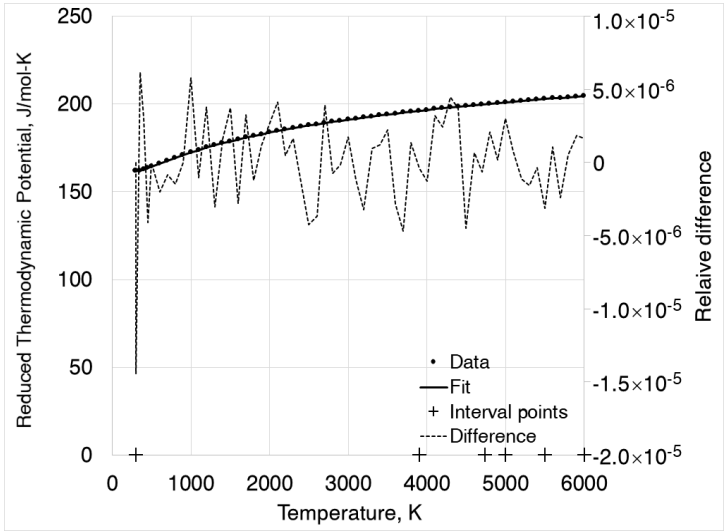
Gallium, Ion (Ga^+) – Ga1+(g)

MW = 69.72 g/mol – $\Delta_f H^o$(298.15 K) = 856.648 kJ/mol – S^o(298.15 K) = 161.79 J/mol-K – p^o = 0.1 MPa



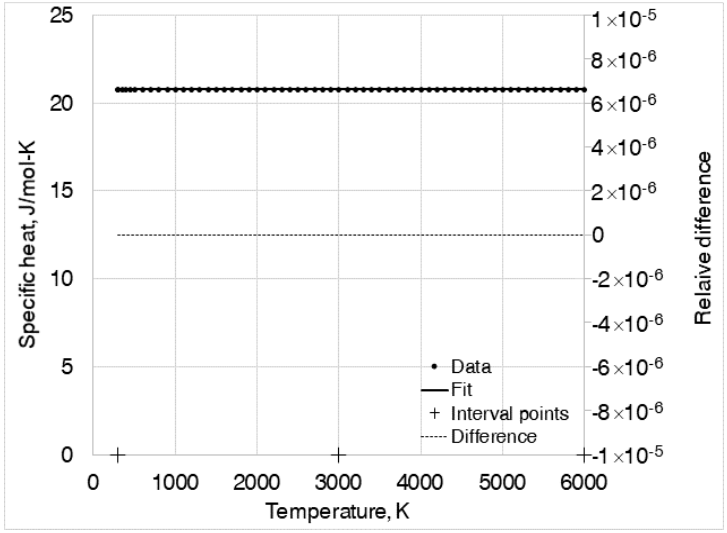


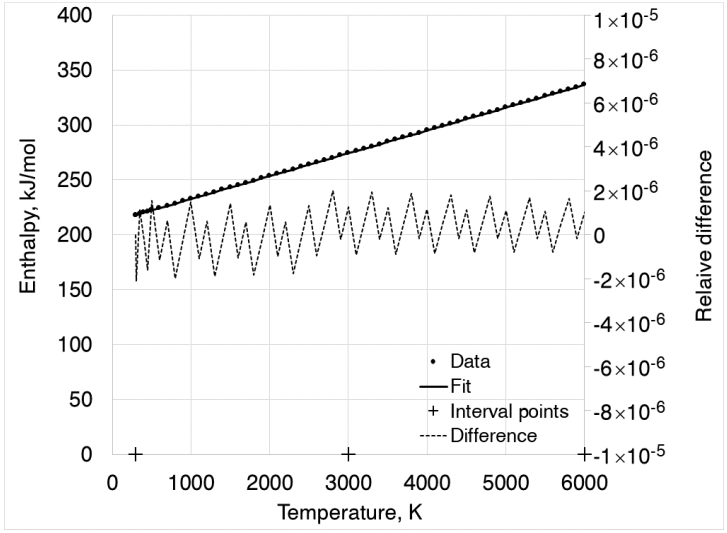


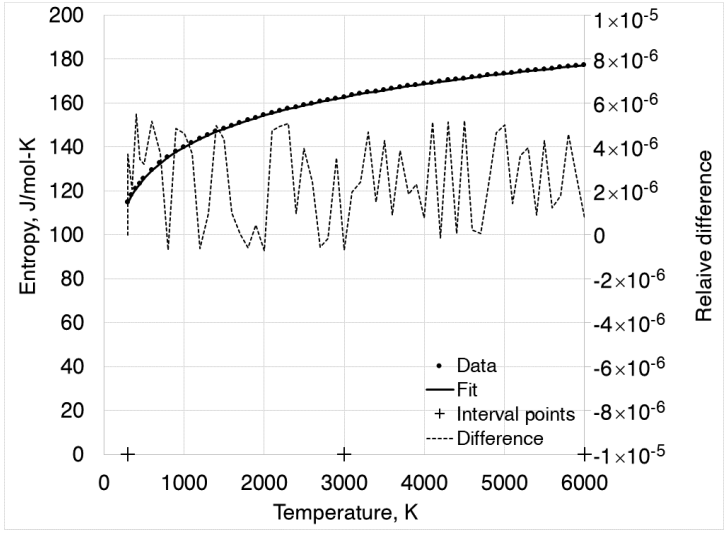


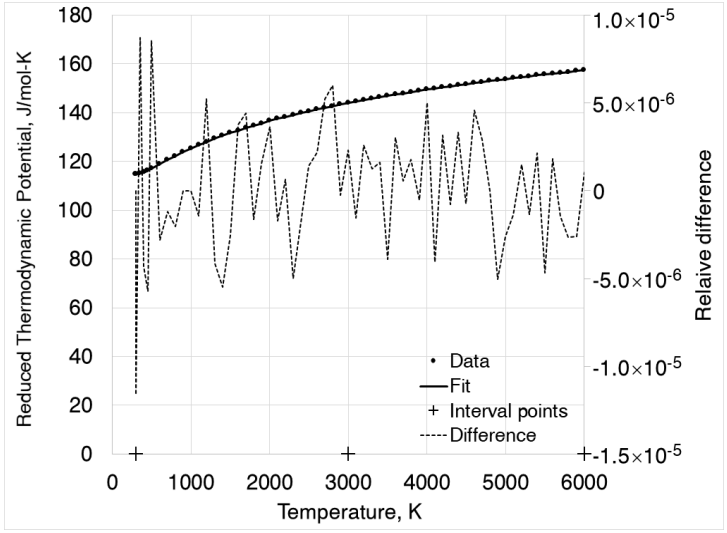
Hydrogen (H) – H1(g)

MW = 1.00794 g/mol – $\Delta_f H^o$(298.15 K) = 217.999 kJ/mol – S^o(298.15 K) = 114.716 J/mol-K – p^o = 0.1 MPa



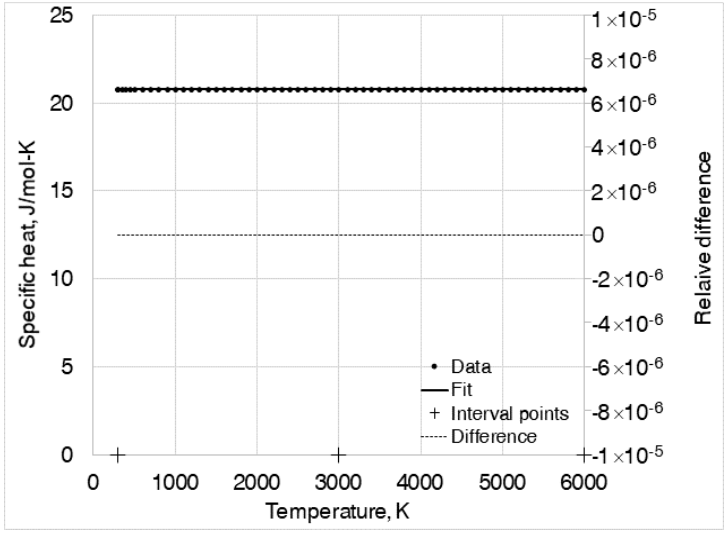


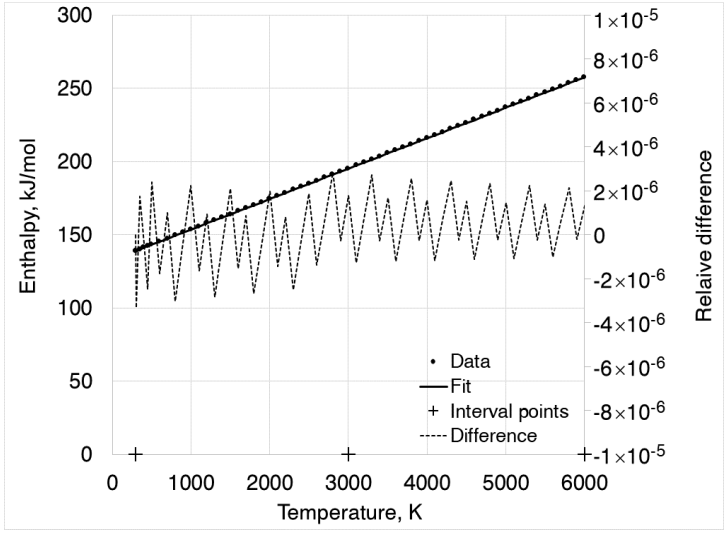


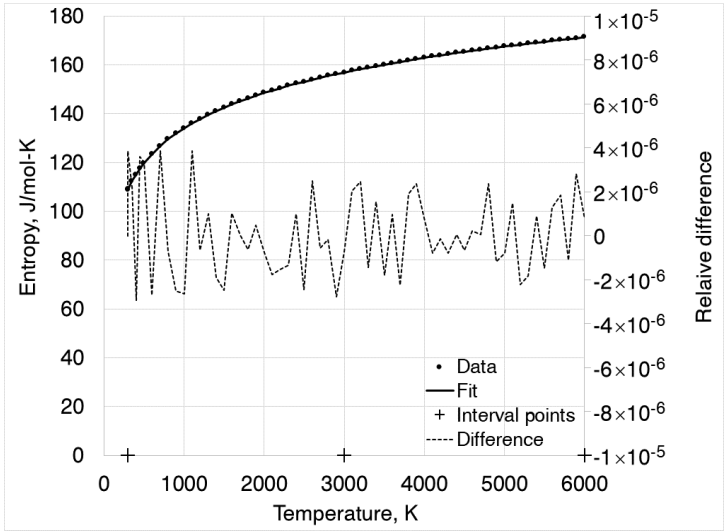


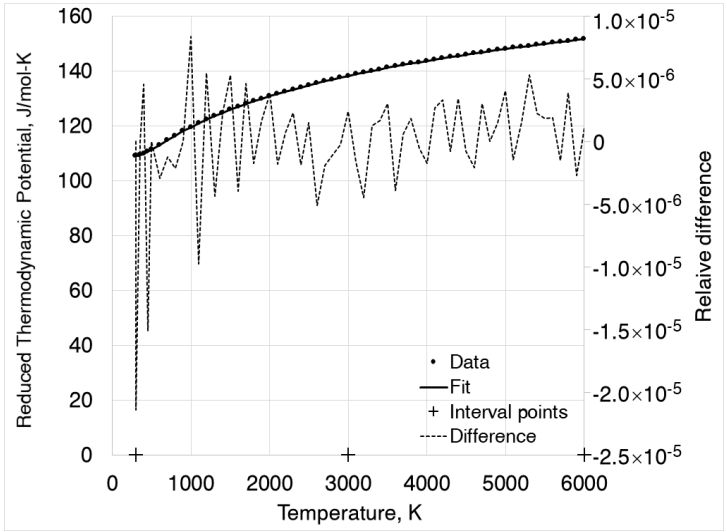
Hydrogen, Ion (H^-) – H1-(g)

MW = 1.00849 g/mol – $\Delta_f H^o$(298.15 K) = 139.032 kJ/mol – S^o(298.15 K) = 108.96 J/mol-K – p^o = 0.1 MPa



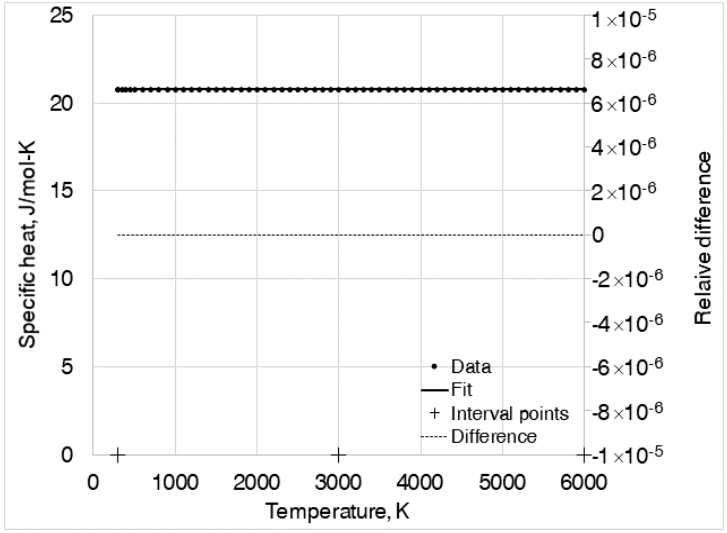


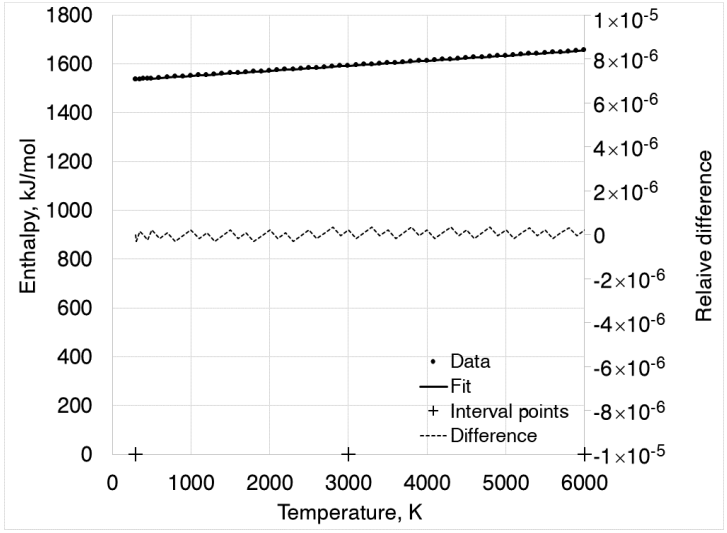


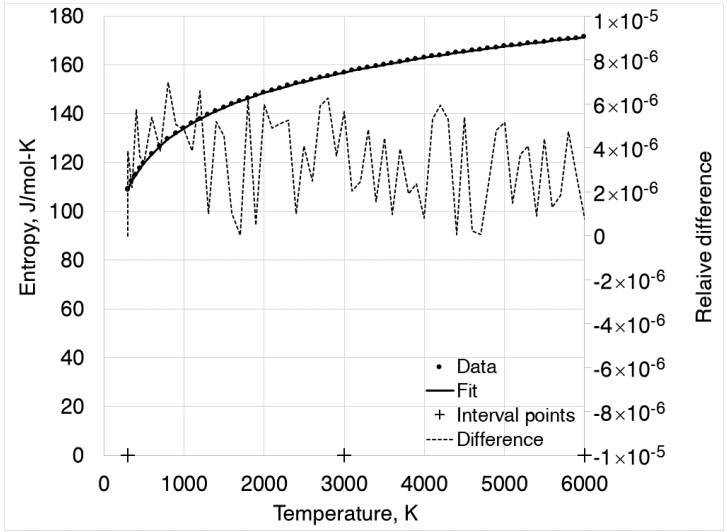


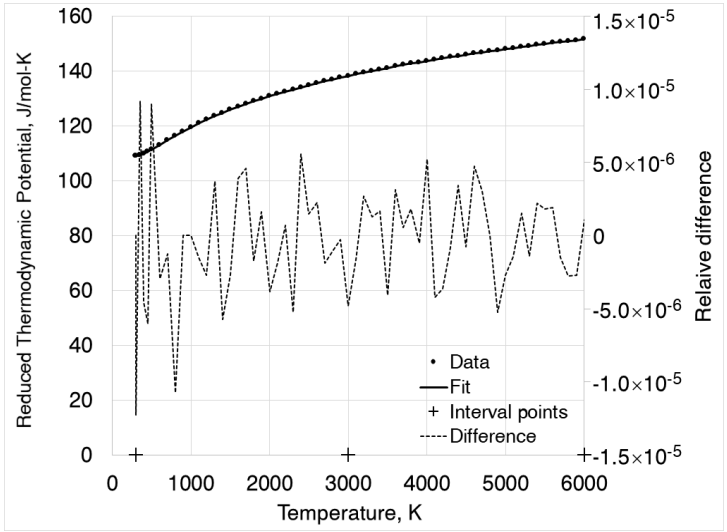
Hydrogen, Ion (H^+) – H1+(g)

MW = 1.00739 g/mol – $\Delta_f H^o$(298.15 K) = 1536.246 kJ/mol – S^o(298.15 K) = 108.946 J/mol-K – p^o = 0.1 MPa



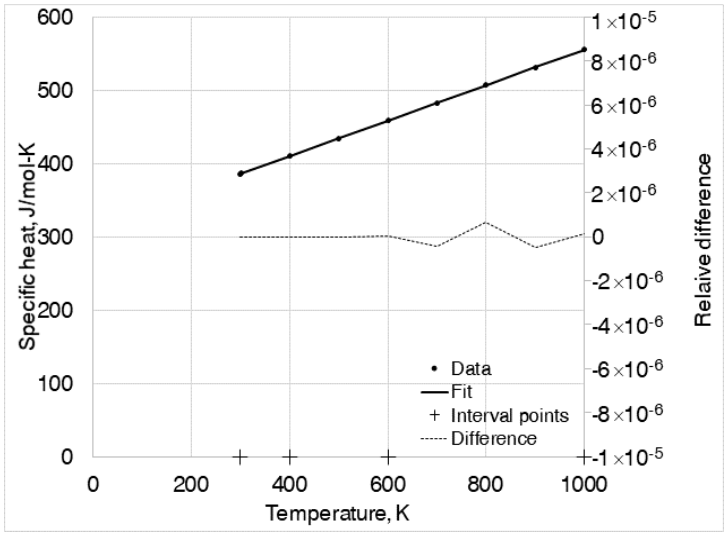


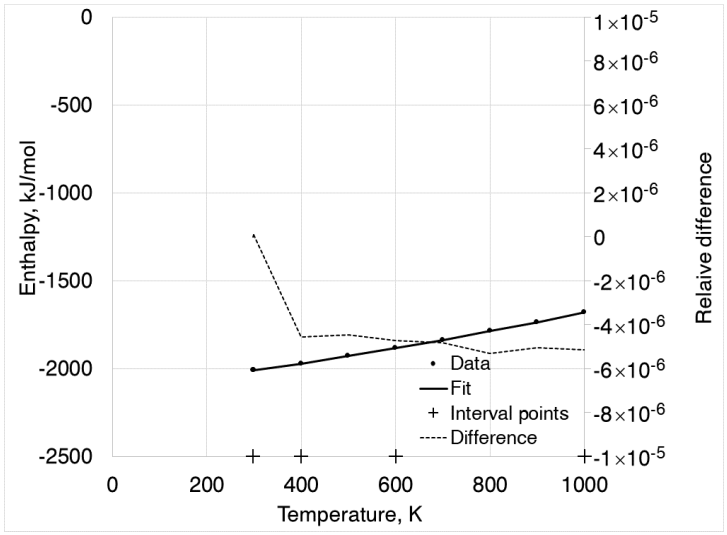


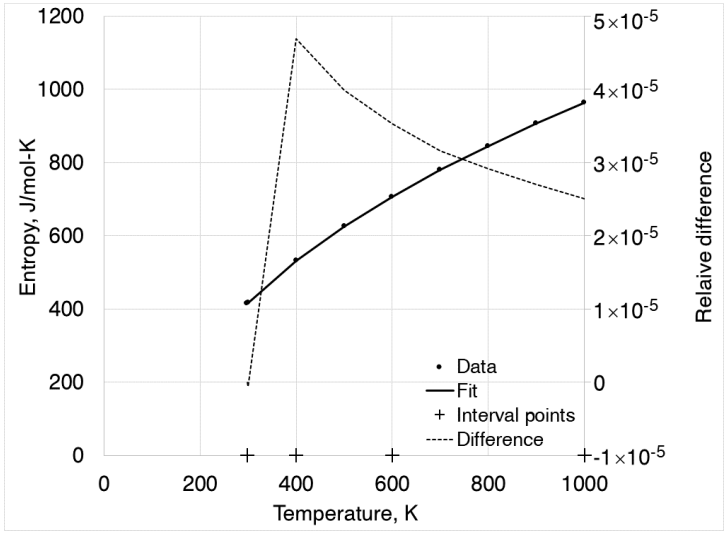


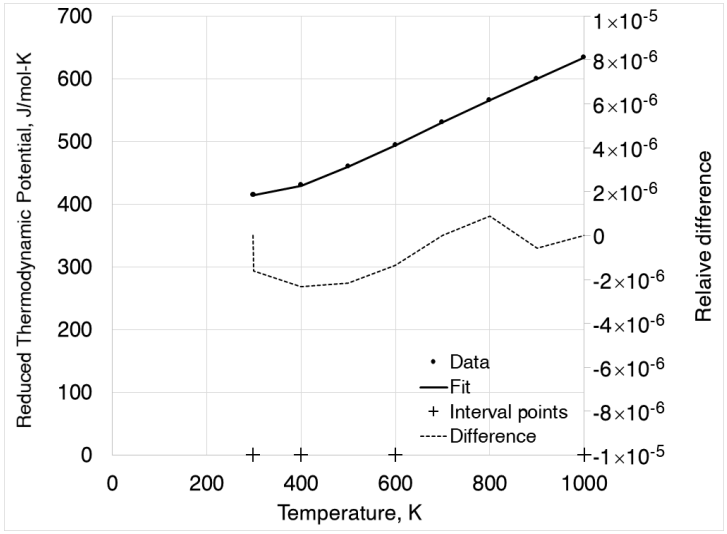
Sulfuric Acid, Tetrahydrate ($H_2SO_4.4H_2O$) – H10O8S1(cr,l)

MW = 170.13 g/mol – $\Delta_f H^o$(298.15 K) = -2011.199 kJ/mol – S^o(298.15 K) = 414.529 J/mol-K – p^o = 0.1 MPa



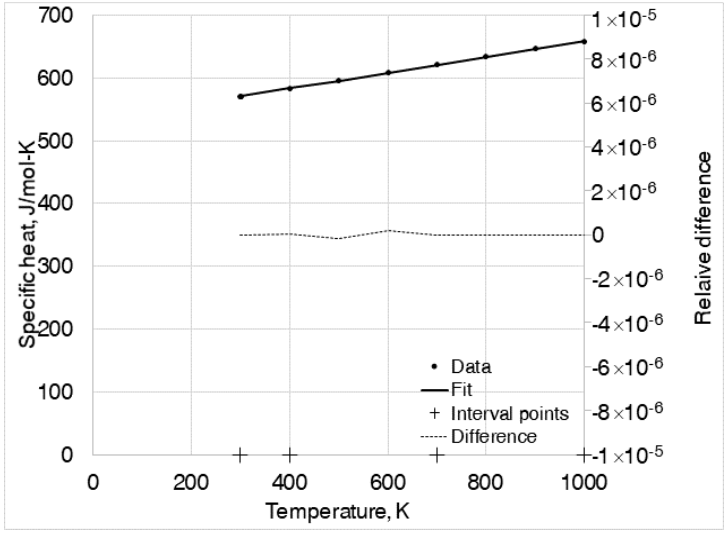


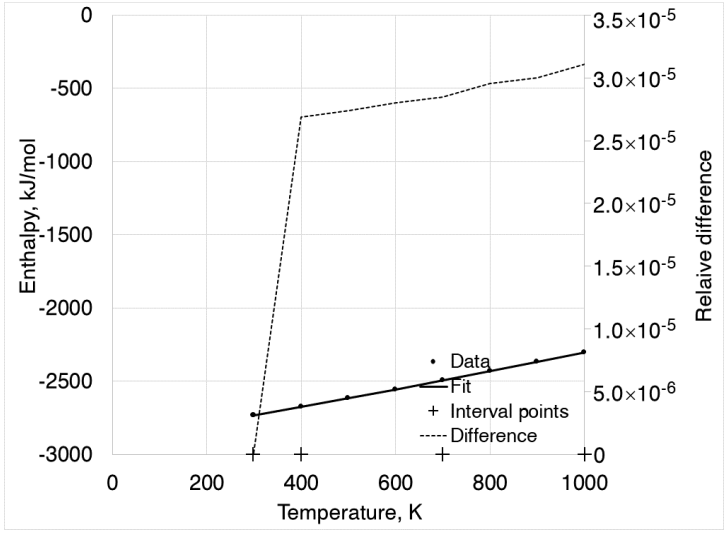


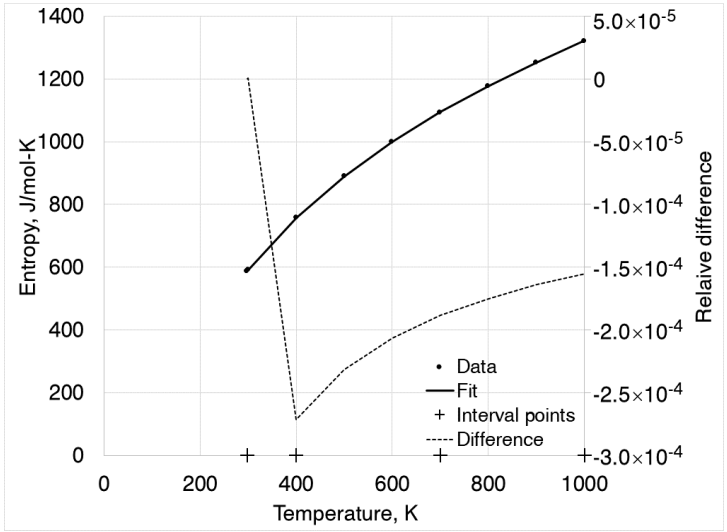


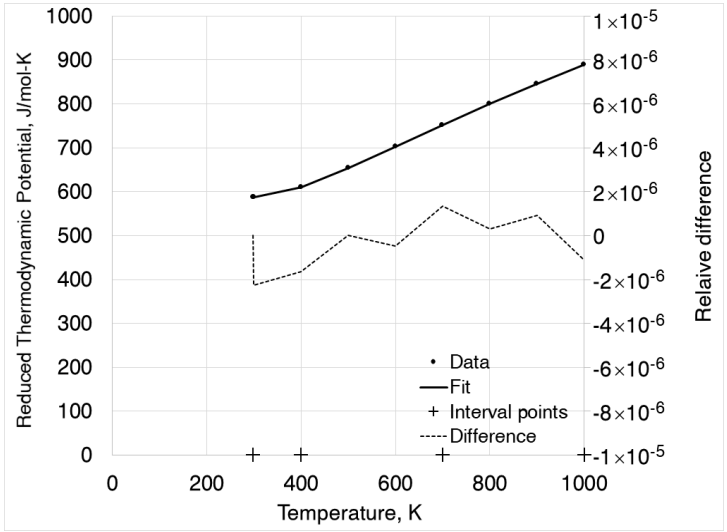
Sulfuric Acid, Hemihexahydrate ($H_2SO_4.6.5H_2O$) – H15O10.5S1(cr,l)

MW = 215.17 g/mol – $\Delta_f H^o$(298.15 K) = -2733.261 kJ/mol – S^o(298.15 K) = 587.819 J/mol-K – p^o = 0.1 MPa



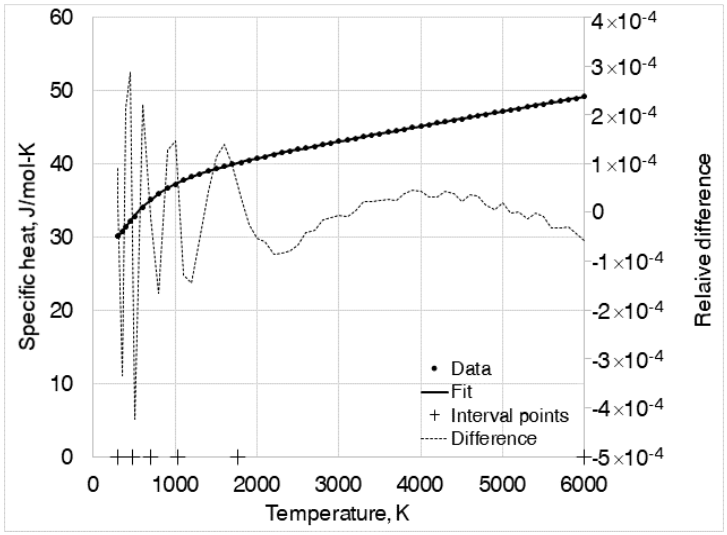


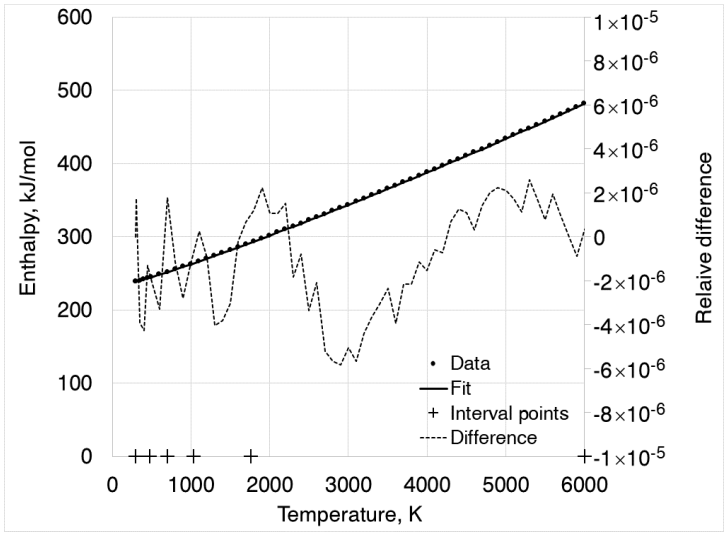


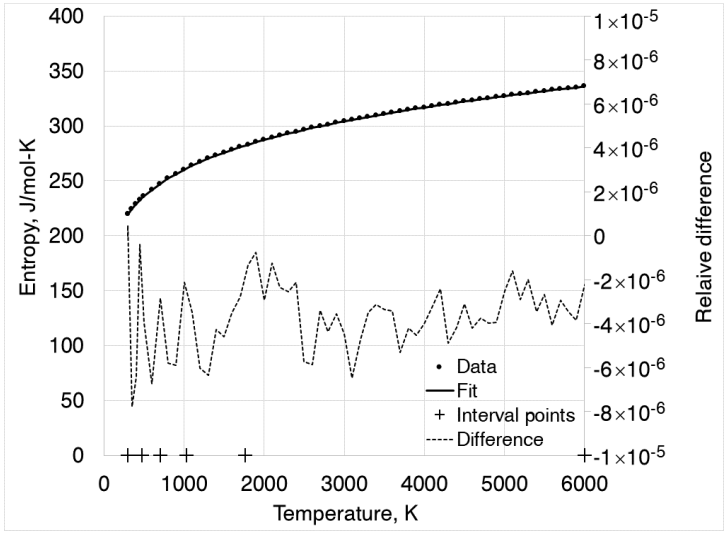


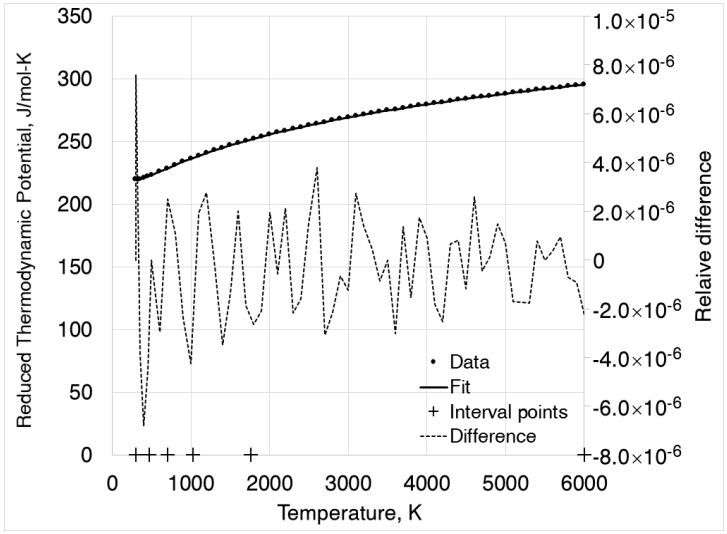
Mercury Hydride (HgH) – H1Hg1(g)

MW = 201.6 g/mol – $\Delta_f H^o$(298.15 K) = 238.488 kJ/mol – S^o(298.15 K) = 219.705 J/mol-K – p^o = 0.1 MPa



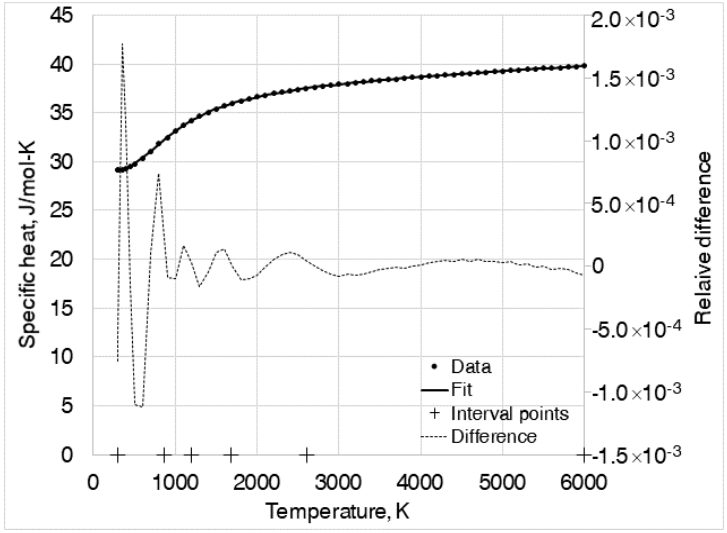


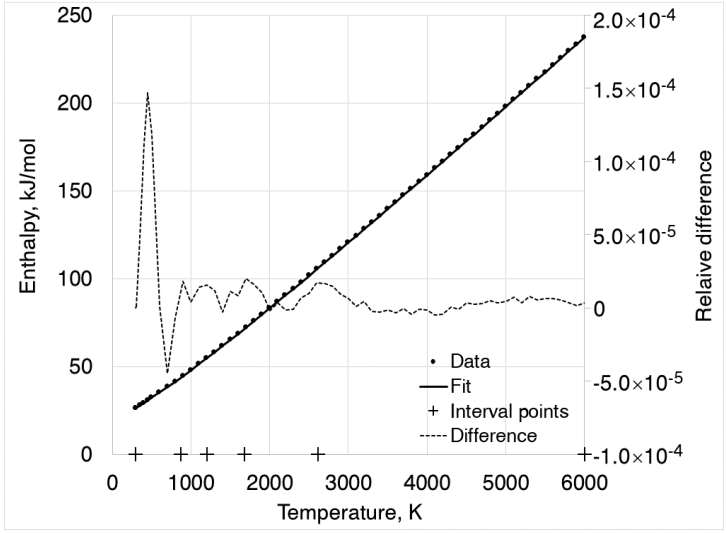


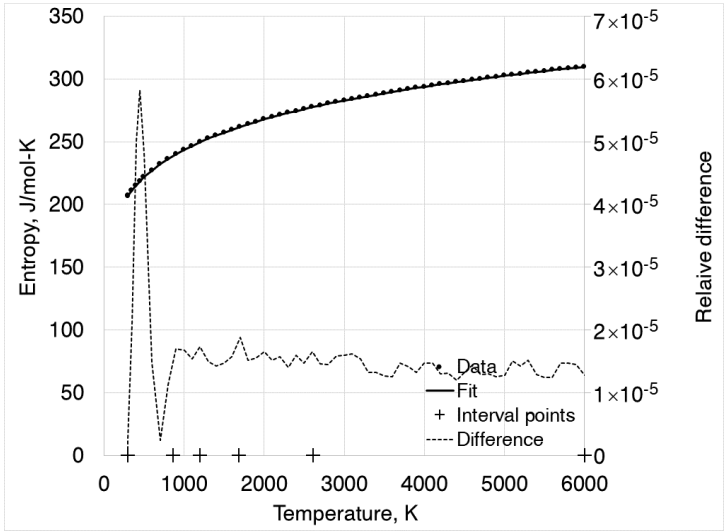


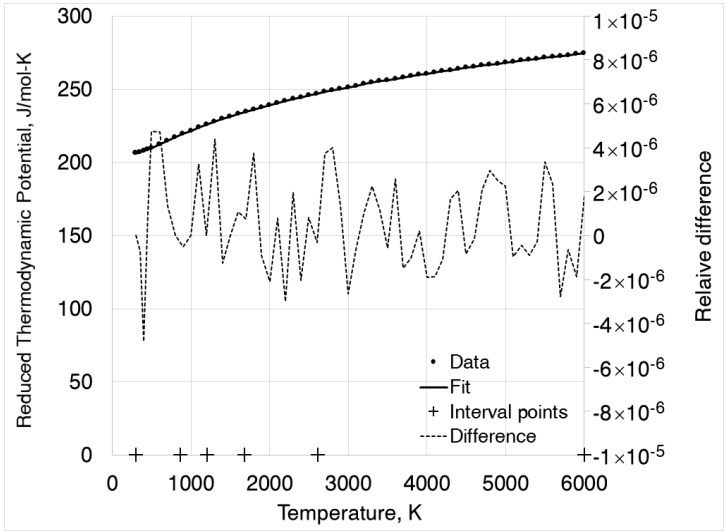
Hydrogen Iodide (HI) – H1I1(g)

MW = 127.9124 g/mol – $\Delta_f H^o$(298.15 K) = 26.359 kJ/mol – S^o(298.15 K) = 206.589 J/mol-K – p^o = 0.1 MPa



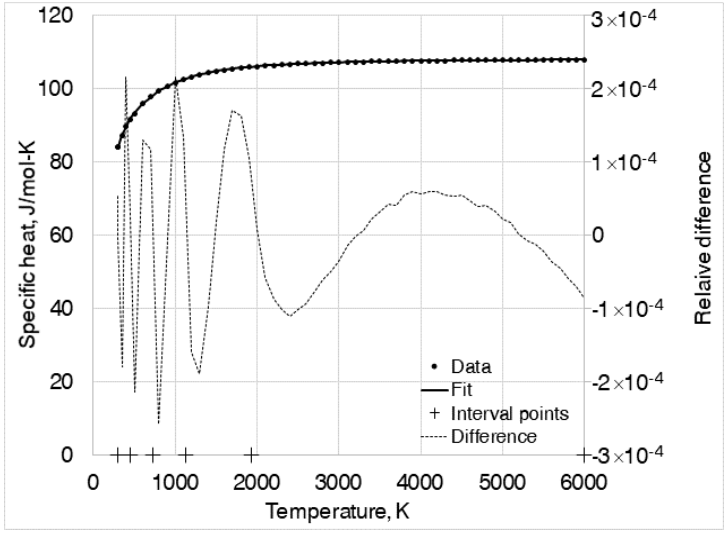


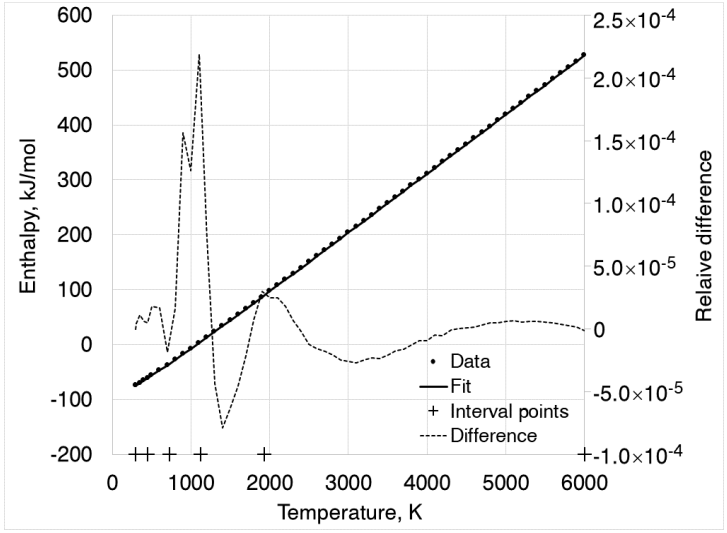


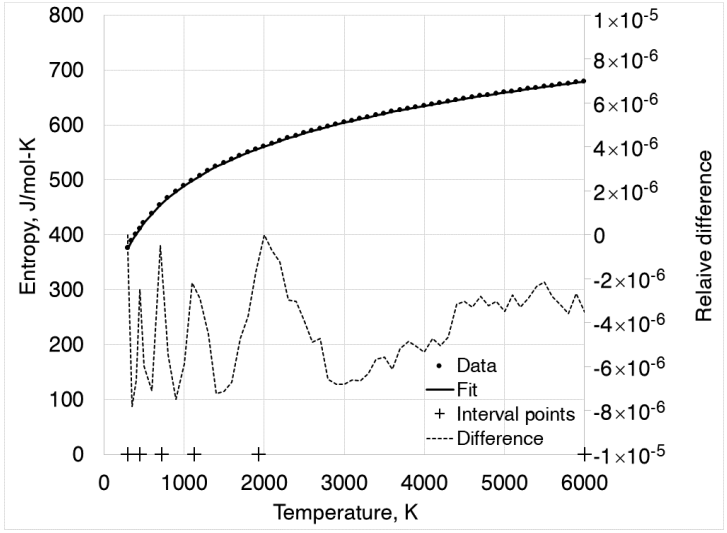


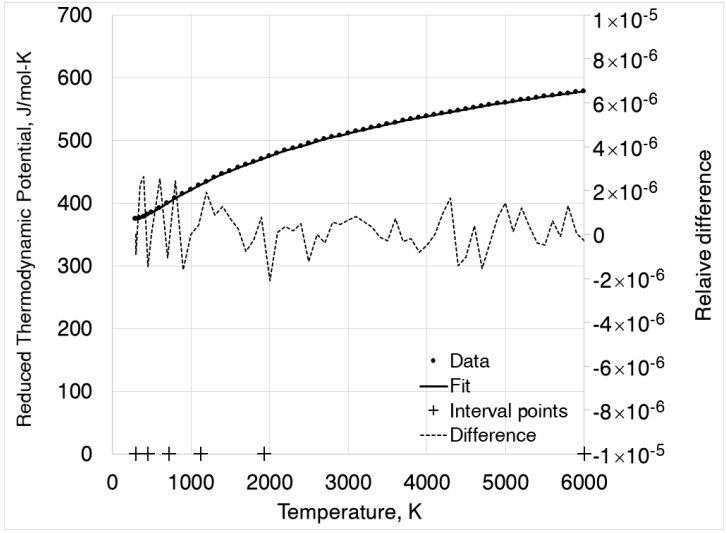
Triiodosilane ($SiHI_3$) – H1I3Si1(g)

MW = 409.8069 g/mol – $\Delta_f H^o$(298.15 K) = -74.475 kJ/mol – S^o(298.15 K) = 375.009 J/mol-K – p^o = 0.1 MPa



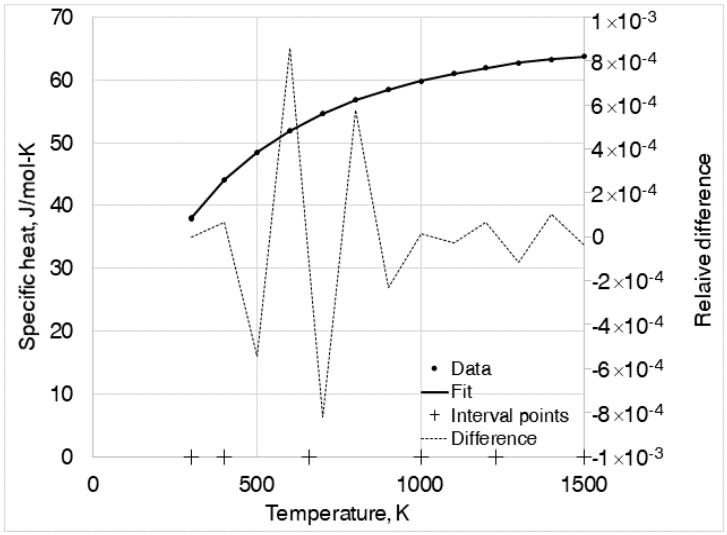


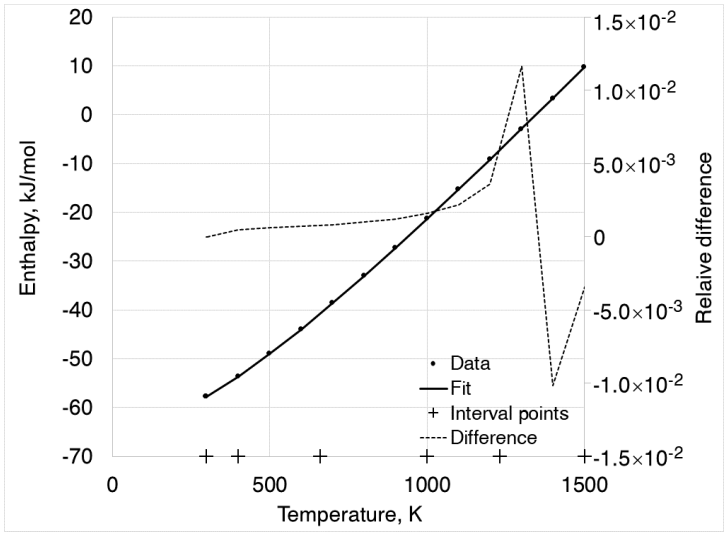


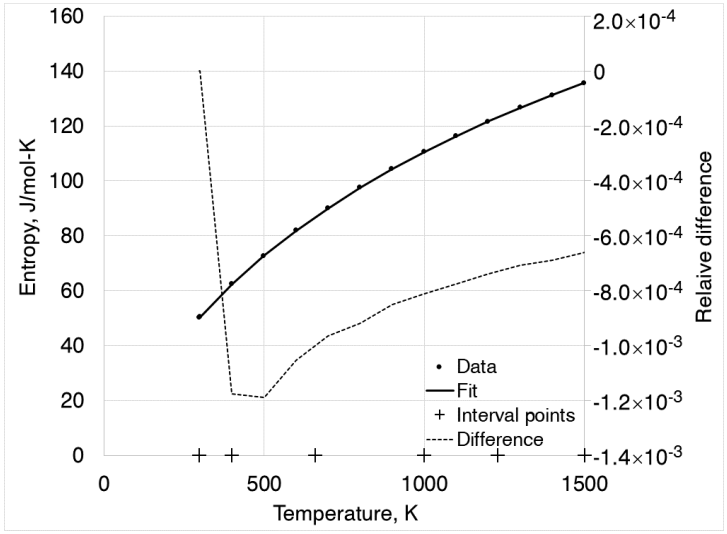


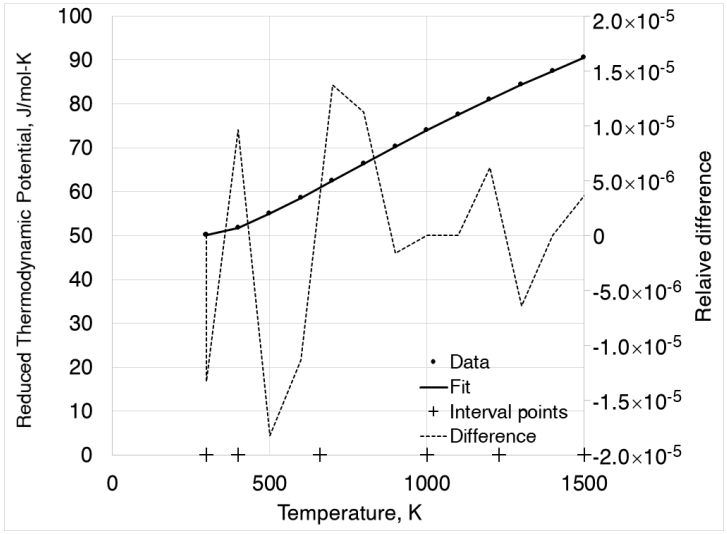
Potassium Hydride (KH) – H1K1(cr)

MW = 40.1062 g/mol – $\Delta_f H^o$(298.15 K) = -57.82 kJ/mol – S^o(298.15 K) = 50.208 J/mol-K – p^o = 0.1 MPa



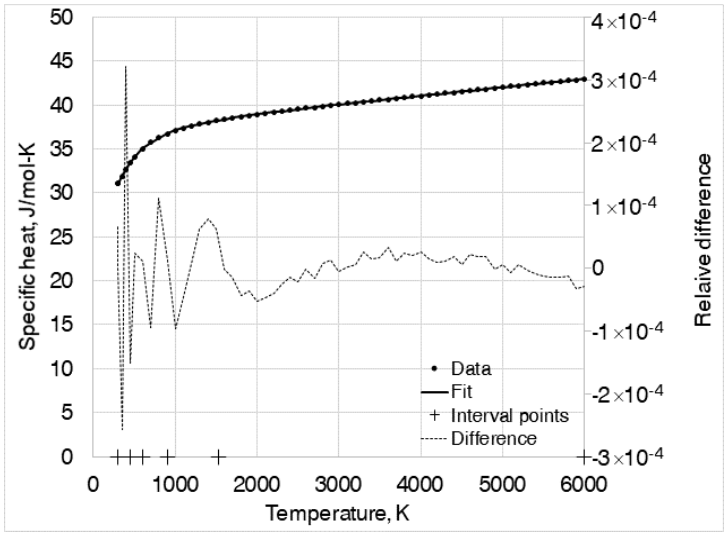


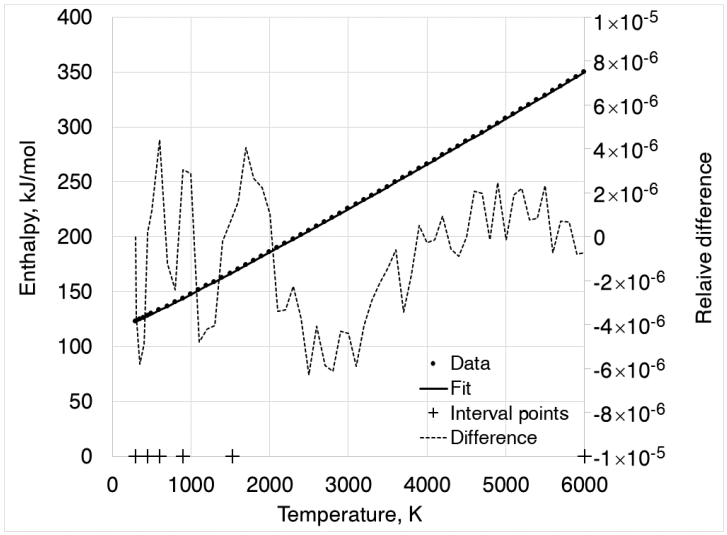


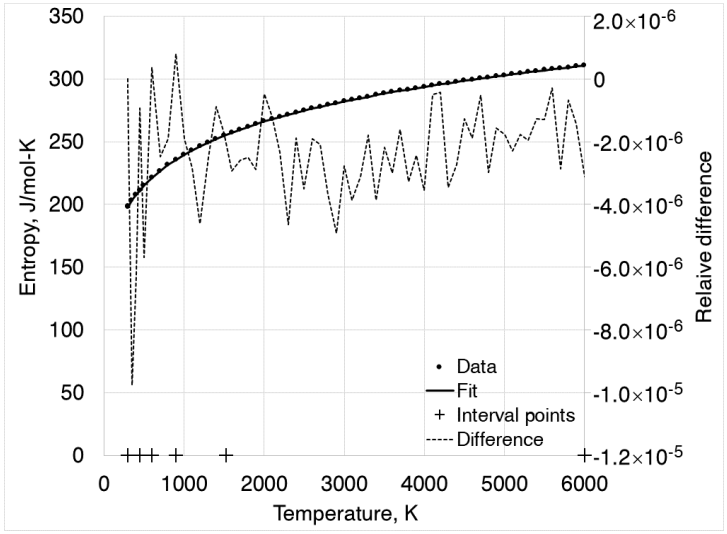


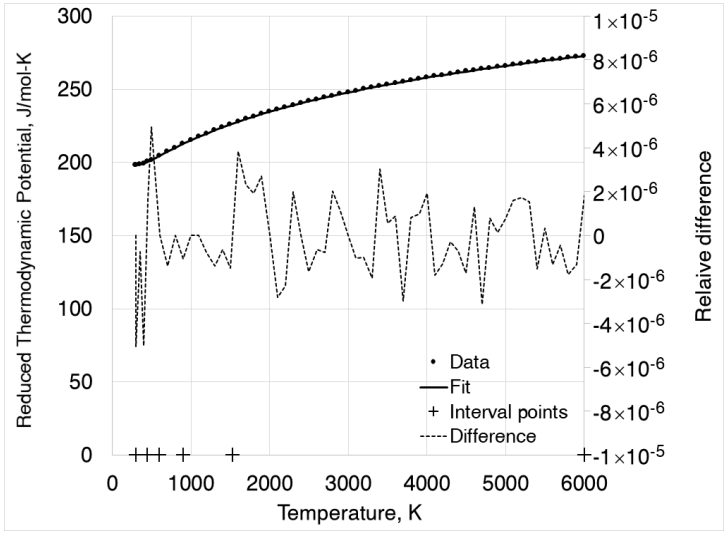
Potassium Hydride (KH) – H1K1(g)

MW = 40.1062 g/mol – $\Delta_f H^o$(298.15 K) = 123.01 kJ/mol – S^o(298.15 K) = 198.024 J/mol-K – p^o = 0.1 MPa



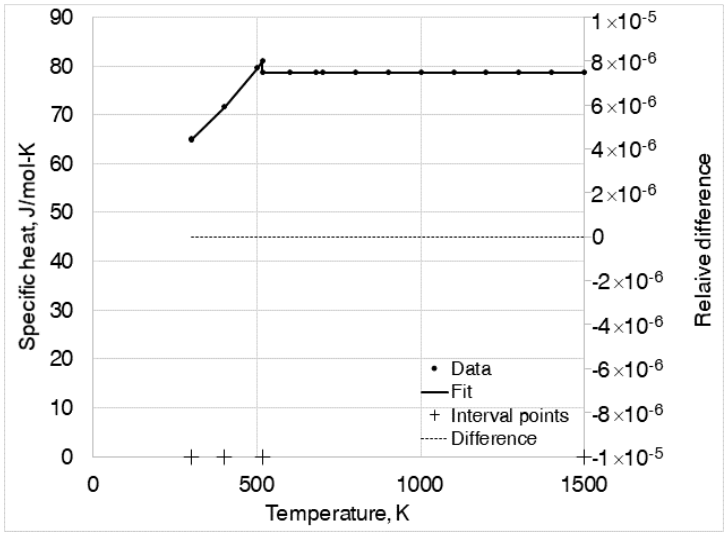


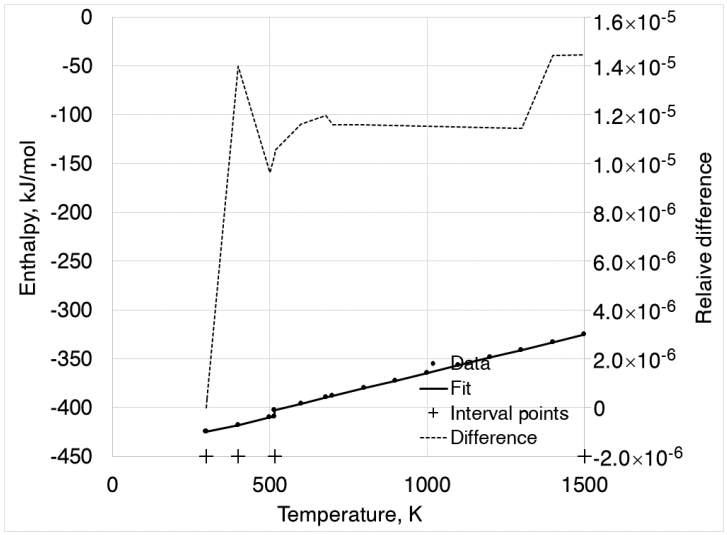


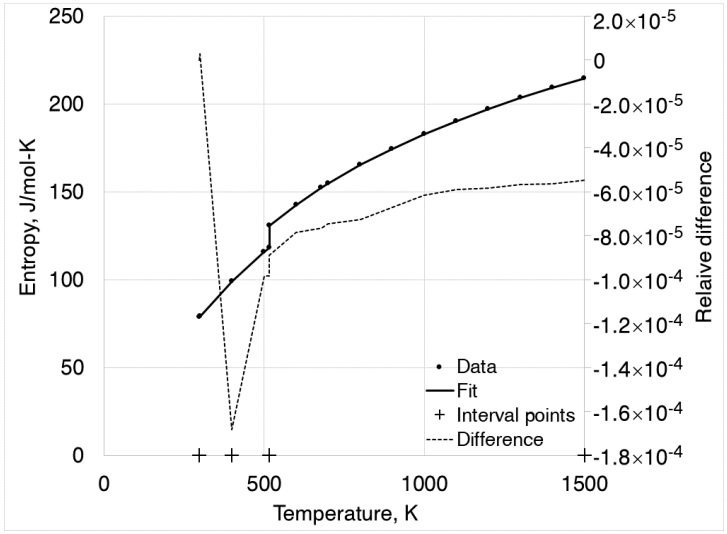


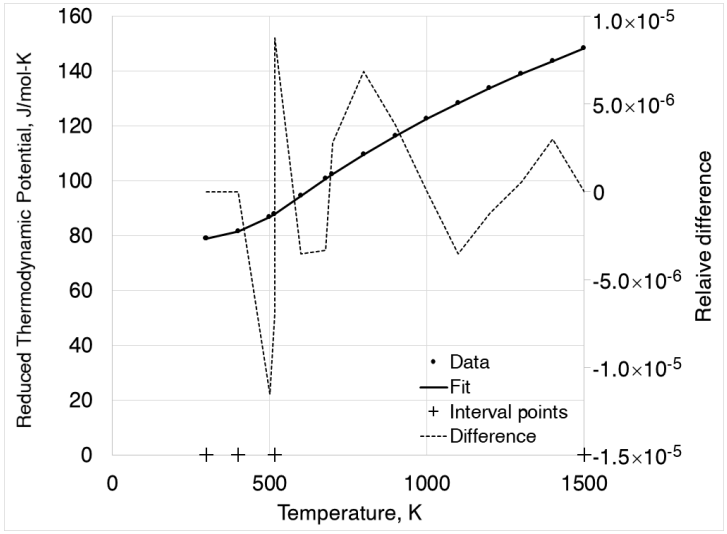
Potassium Hydroxide (KOH) – H1K1O1(cr)

MW = 56.1056 g/mol – $\Delta_f H^o$(298.15 K) = -424.718 kJ/mol – S^o(298.15 K) = 78.907 J/mol-K – p^o = 0.1 MPa



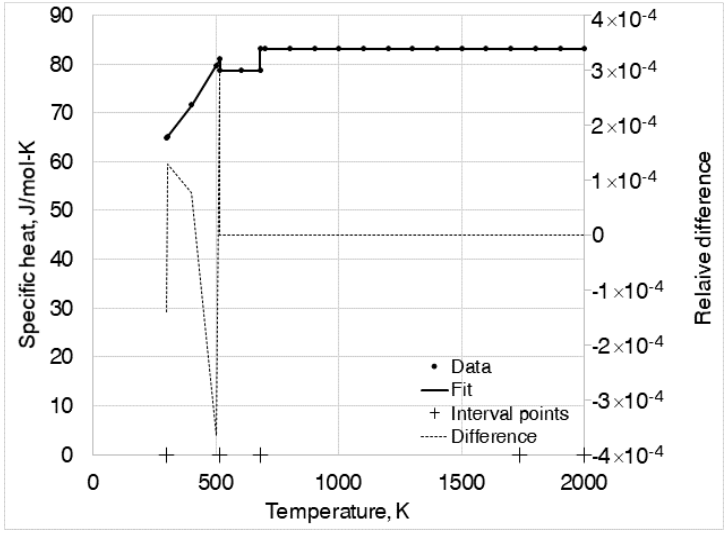


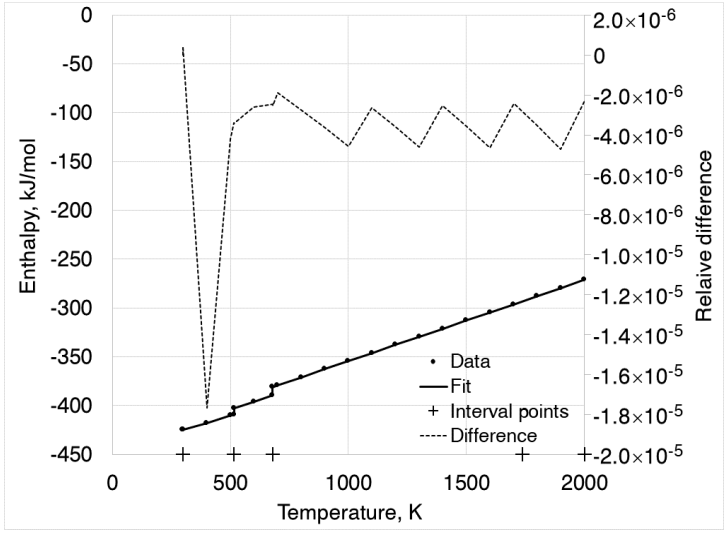


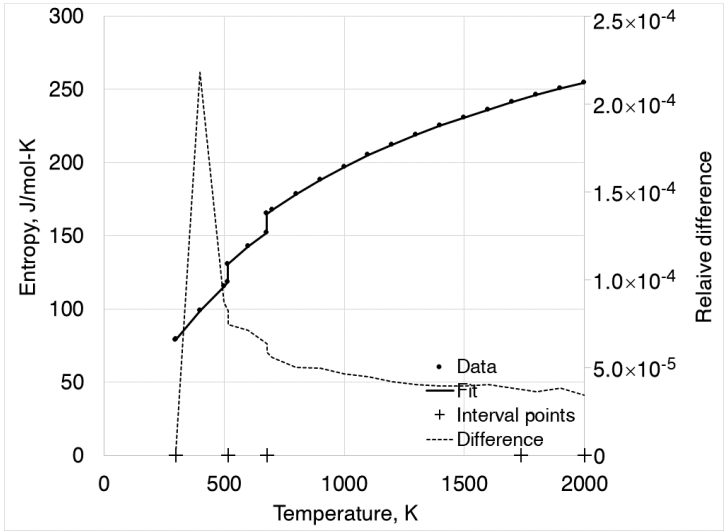


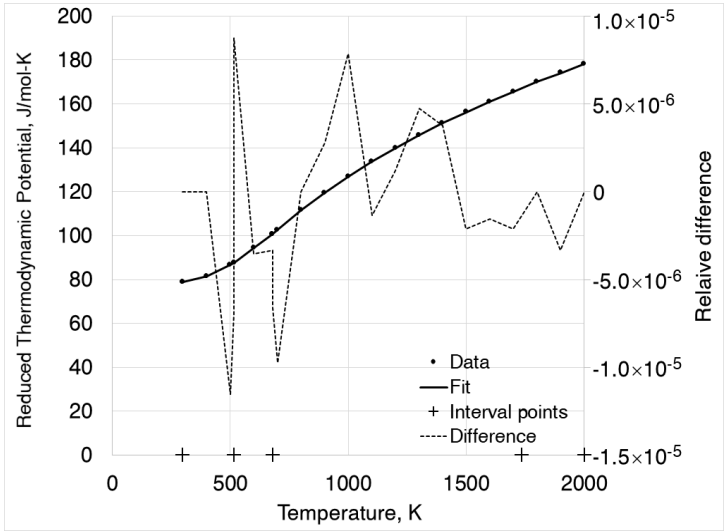
Potassium Hydroxide (KOH) – H1K1O1(cr,l)

MW = 56.1056 g/mol – $\Delta_f H^o$(298.15 K) = -424.718 kJ/mol – S^o(298.15 K) = 78.907 J/mol-K – p^o = 0.1 MPa



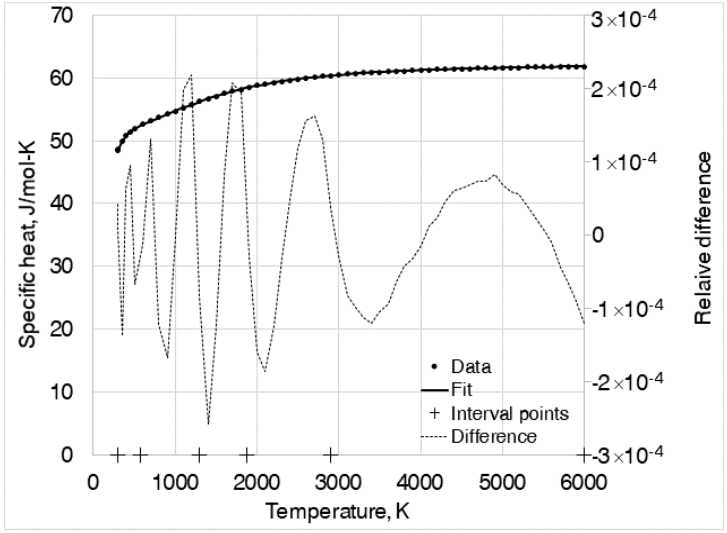


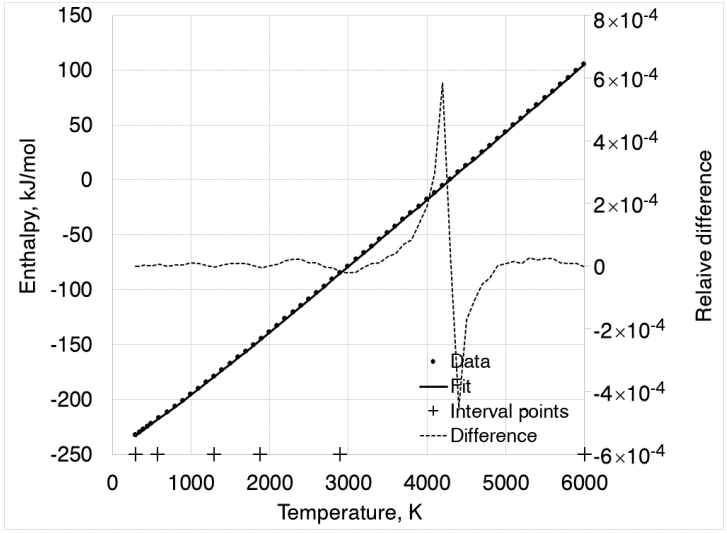


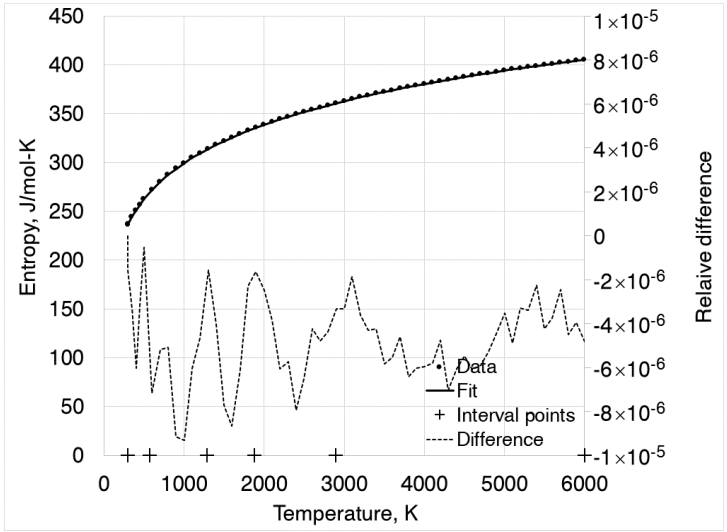


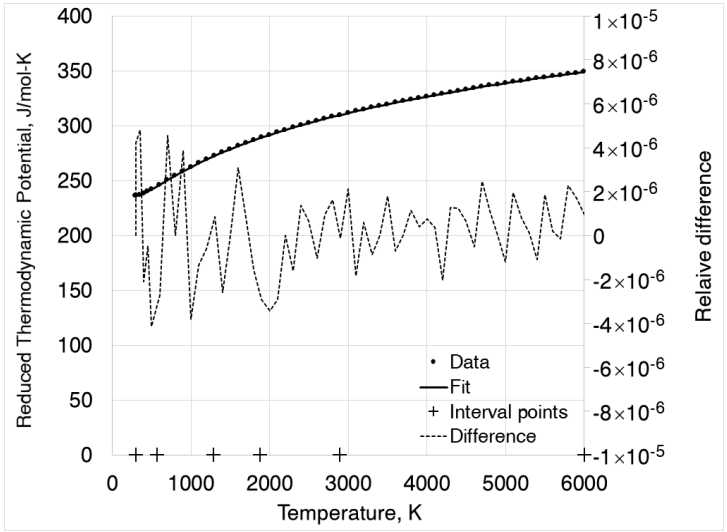
Potassium Hydroxide (KOH) – H1K1O1(g)

MW = 56.1056 g/mol – $\Delta_f H^o$(298.15 K) = -232.63 kJ/mol – S^o(298.15 K) = 236.379 J/mol-K – p^o = 0.1 MPa



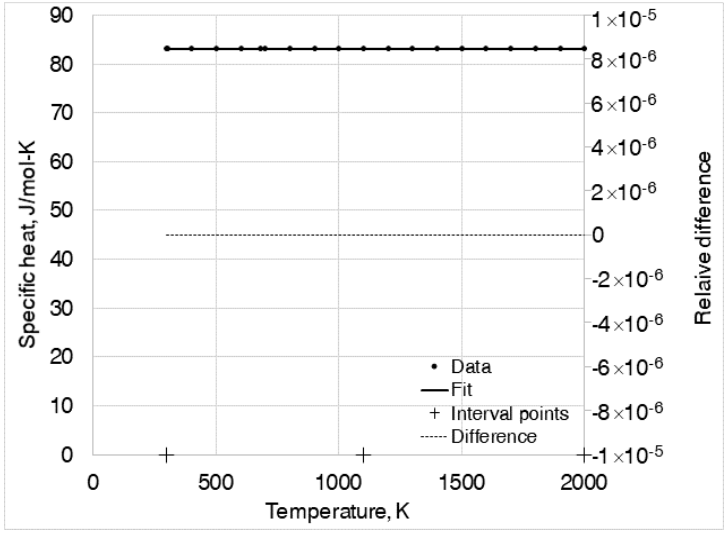


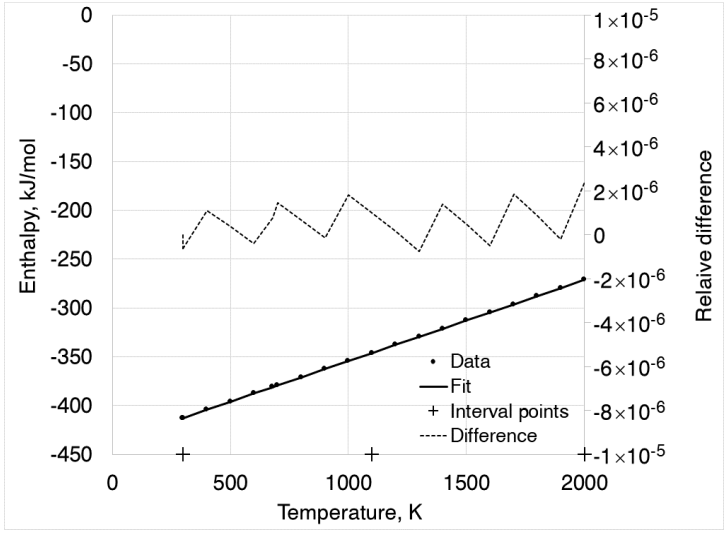


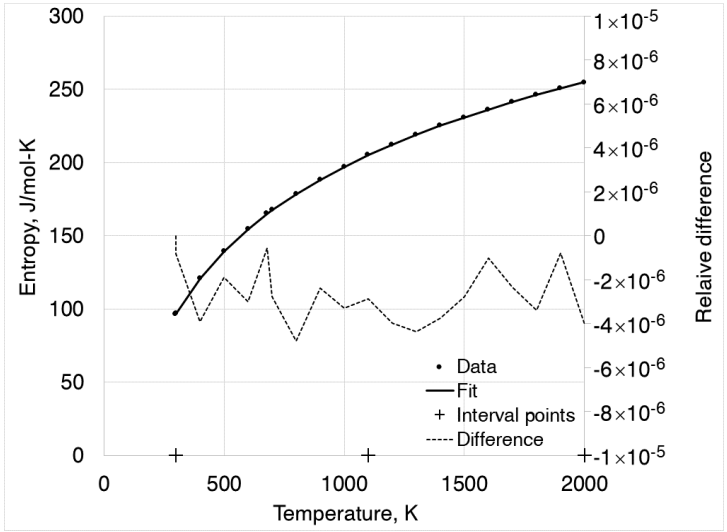


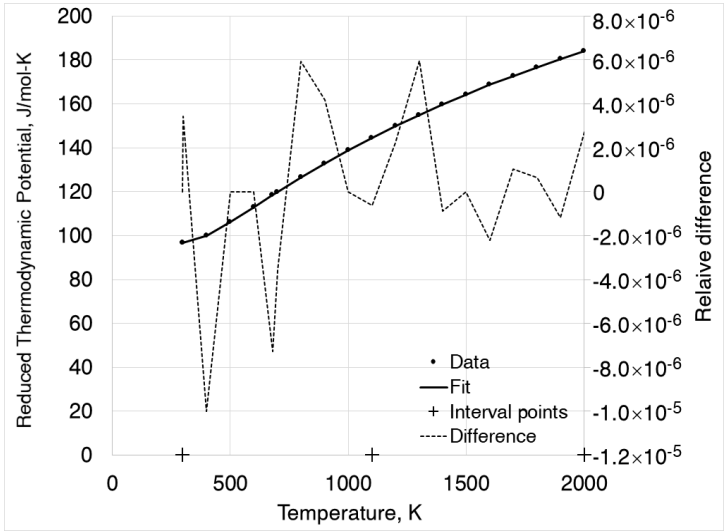
Potassium Hydroxide (KOH) – H1K1O1(l)

MW = 56.1056 g/mol – $\Delta_f H^o$(298.15 K) = -412.705 kJ/mol – S^o(298.15 K) = 96.617 J/mol-K – p^o = 0.1 MPa



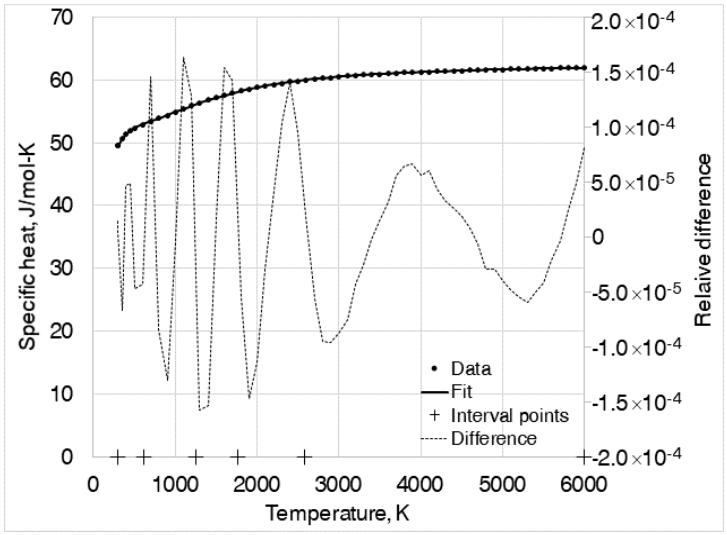


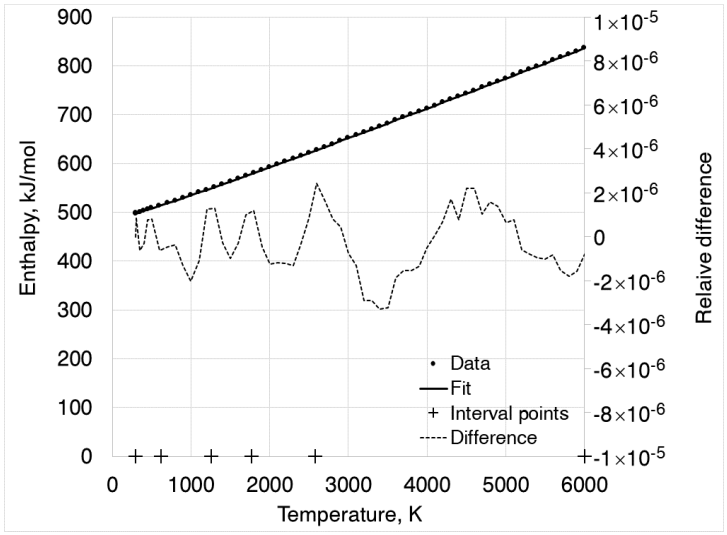


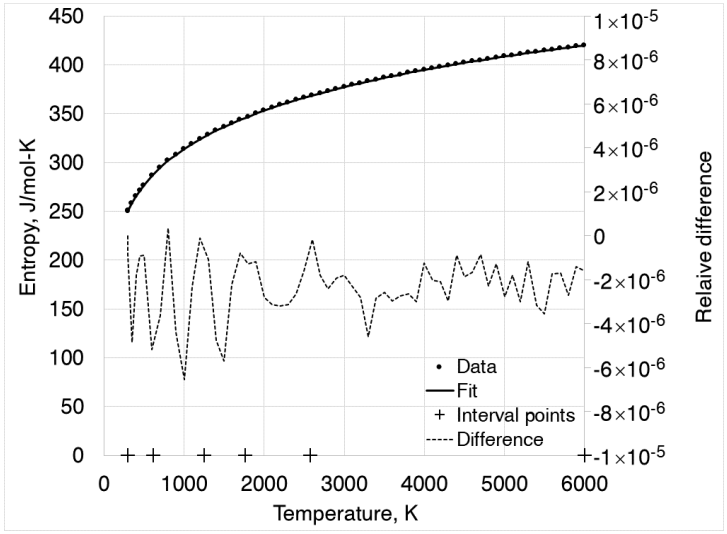


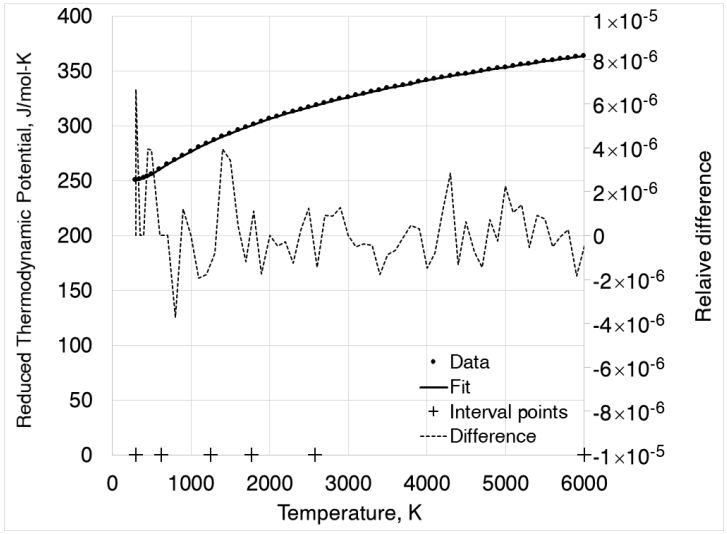
Potassium Hydroxide, Ion (KOH^+) – H1K1O1+(g)

MW = 56.1051 g/mol – $\Delta_f H^o$(298.15 K) = 497.896 kJ/mol – S^o(298.15 K) = 250.331 J/mol-K – p^o = 0.1 MPa



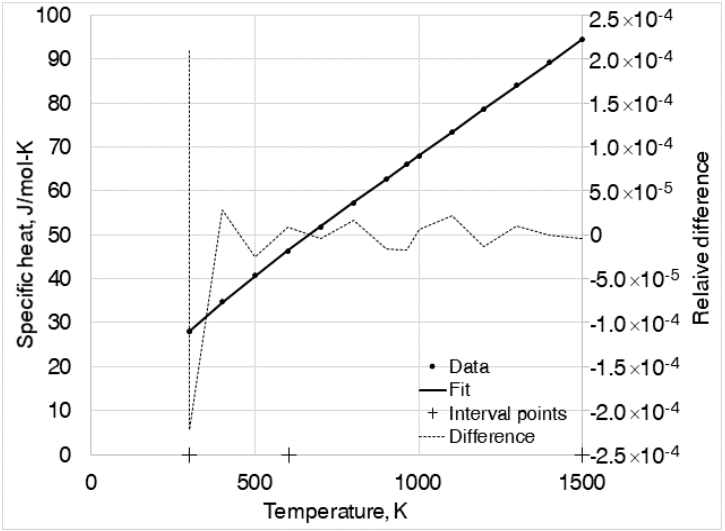


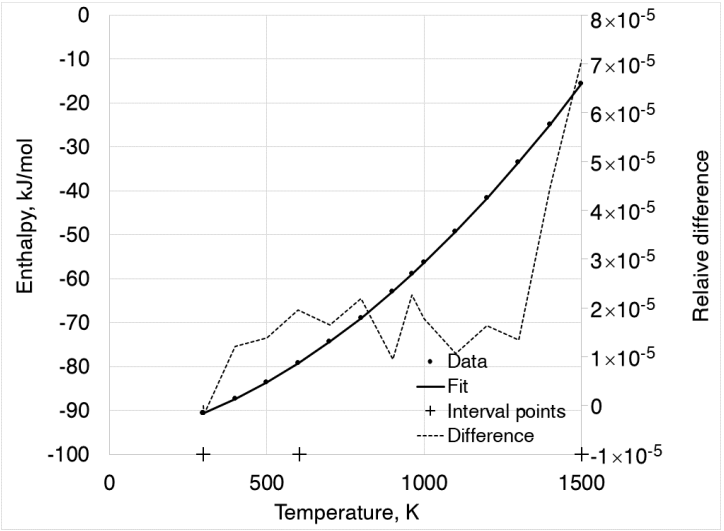


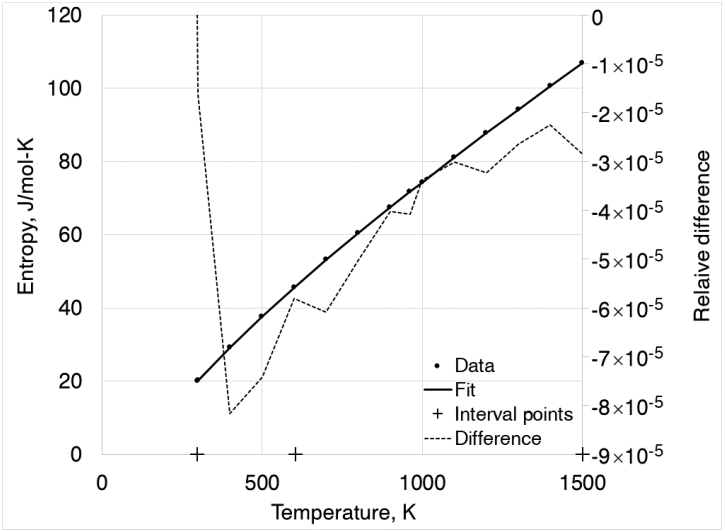


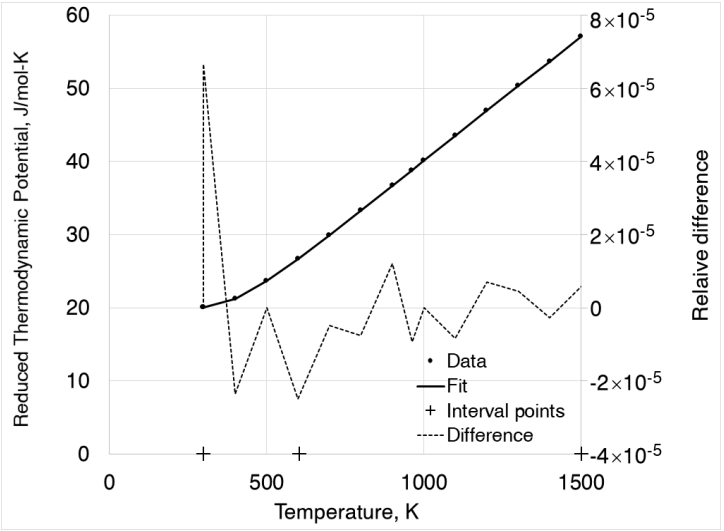
Lithium Hydride (LiH) – H1Li1(cr)

MW = 7.949 g/mol – $\Delta_f H^o$(298.15 K) = -90.625 kJ/mol – S^o(298.15 K) = 20.037 J/mol-K – p^o = 0.1 MPa



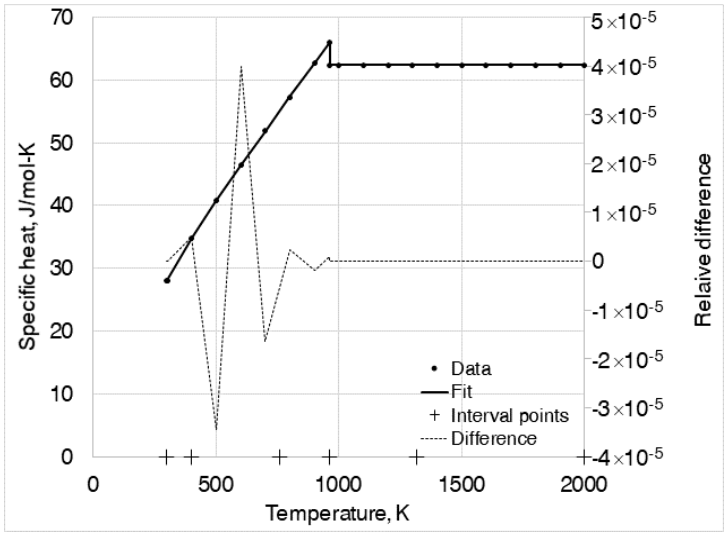


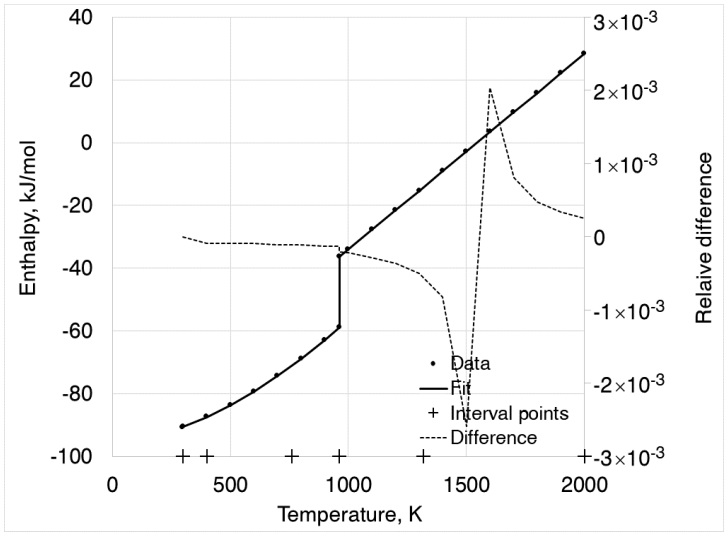


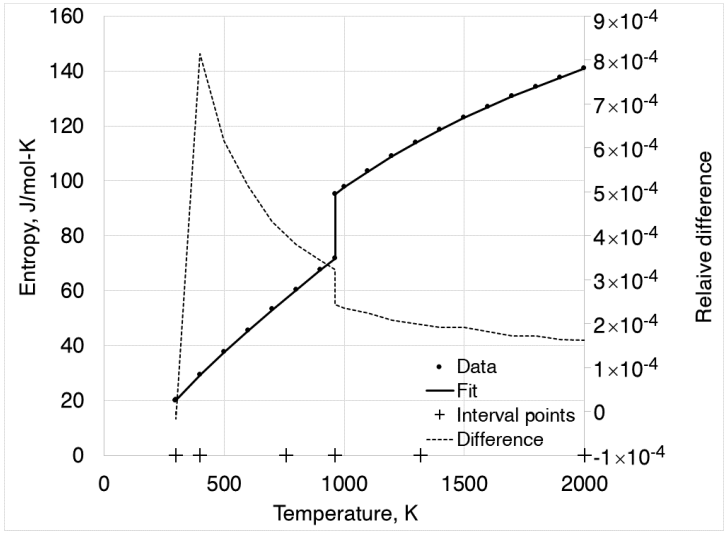


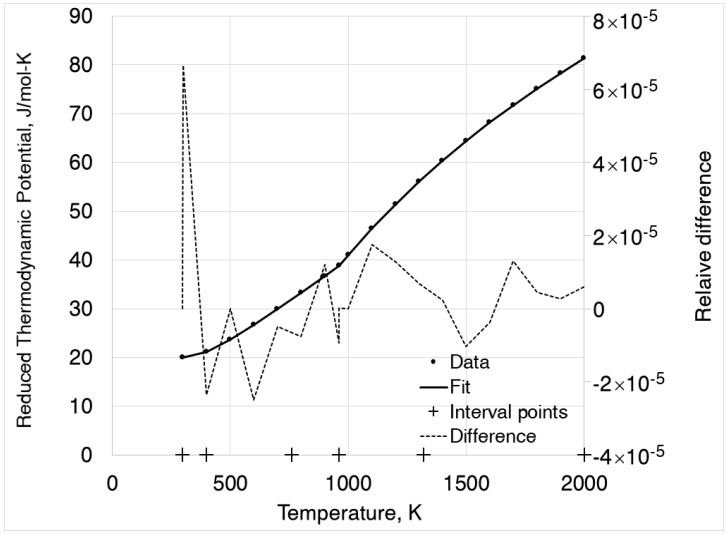
Lithium Hydride (LiH) – H1Li1(cr,l)

MW = 7.949 g/mol – $\Delta_f H^o$(298.15 K) = -90.625 kJ/mol – S^o(298.15 K) = 20.037 J/mol-K – p^o = 0.1 MPa



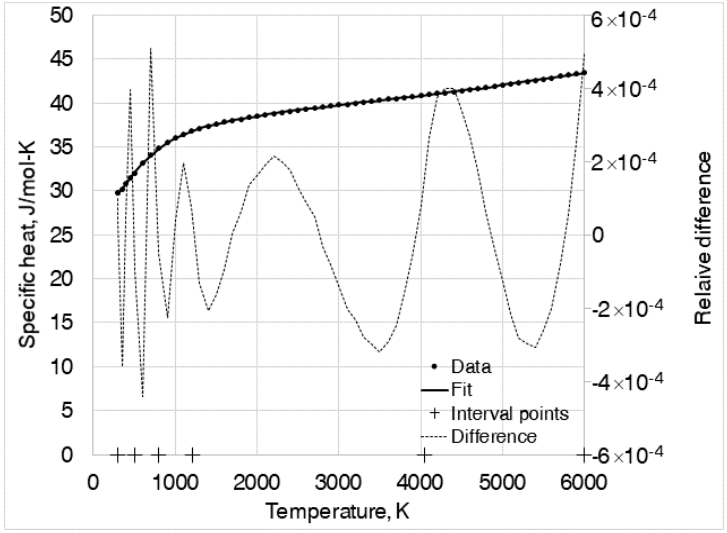


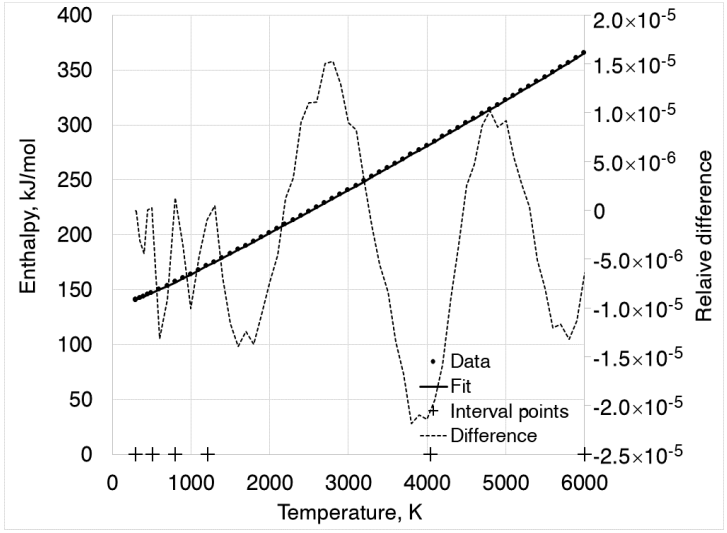


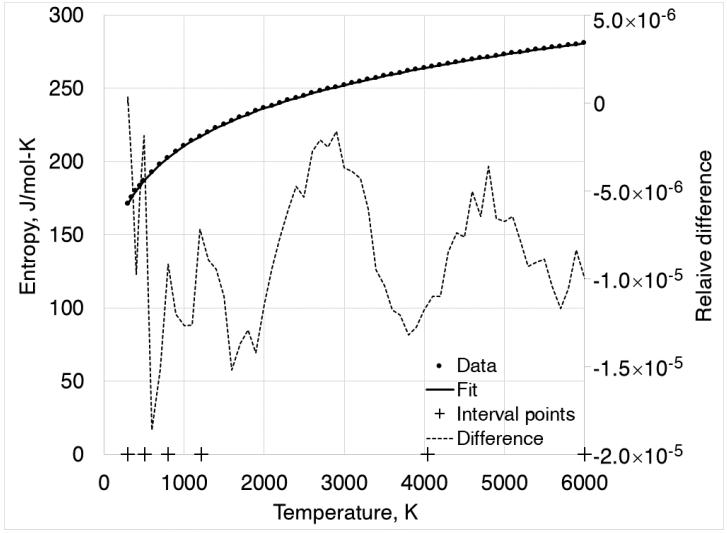


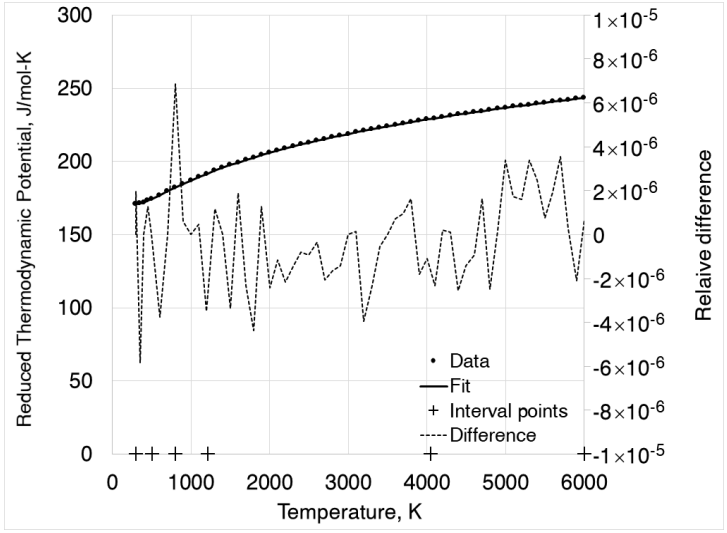
Lithium Hydride (LiH) – H1Li1(g)

MW = 7.949 g/mol – $\Delta_f H^o$(298.15 K) = 140.624 kJ/mol – S^o(298.15 K) = 170.909 J/mol-K – p^o = 0.1 MPa



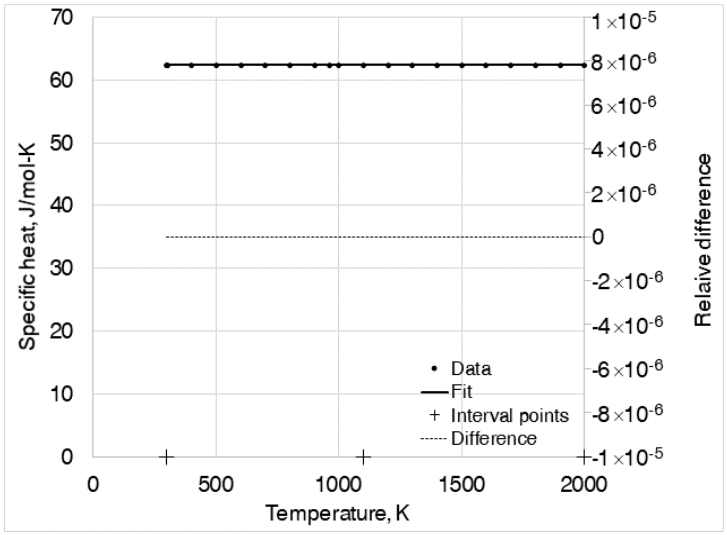


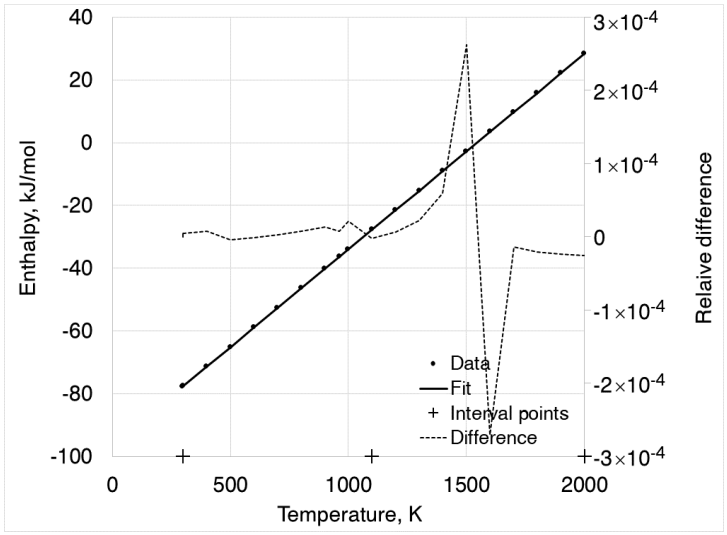


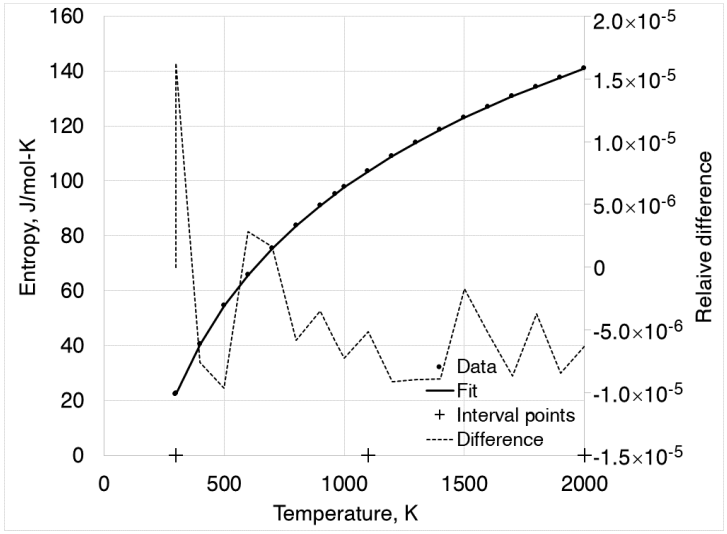


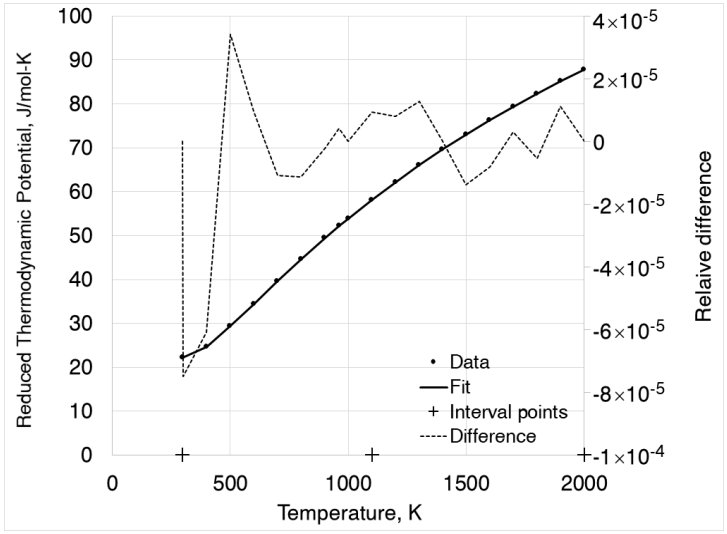
Lithium Hydride (LiH) – H1Li1(l)

MW = 7.949 g/mol – $\Delta_f H^o$(298.15 K) = -77.714 kJ/mol – S^o(298.15 K) = 22.227 J/mol-K – p^o = 0.1 MPa



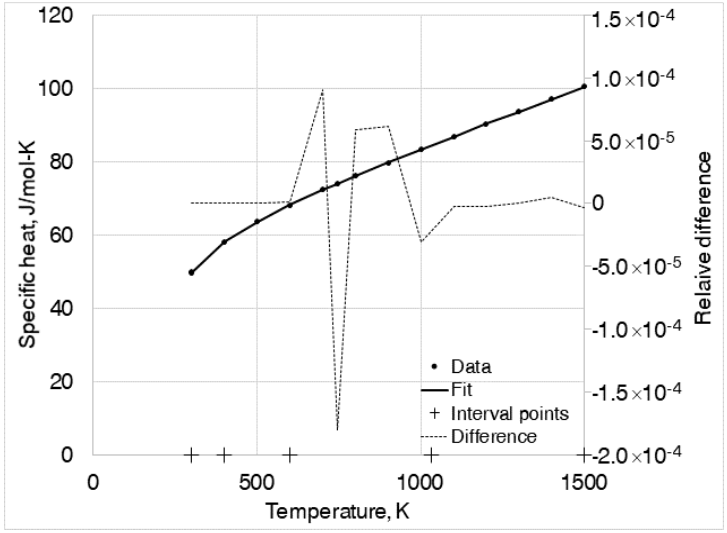


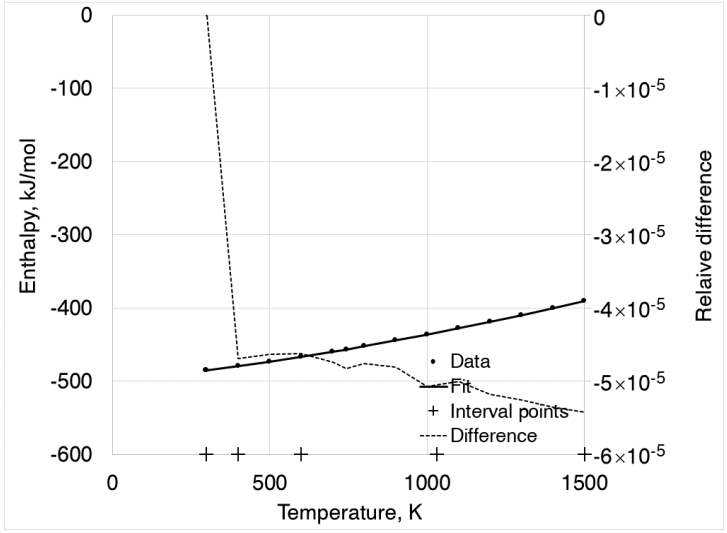


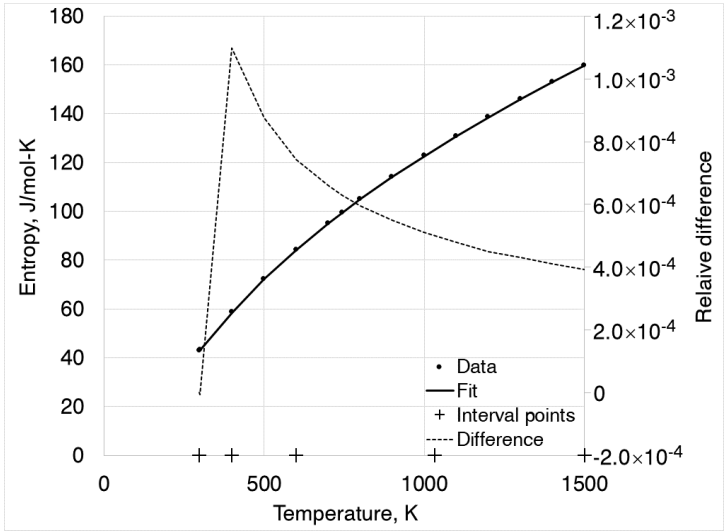


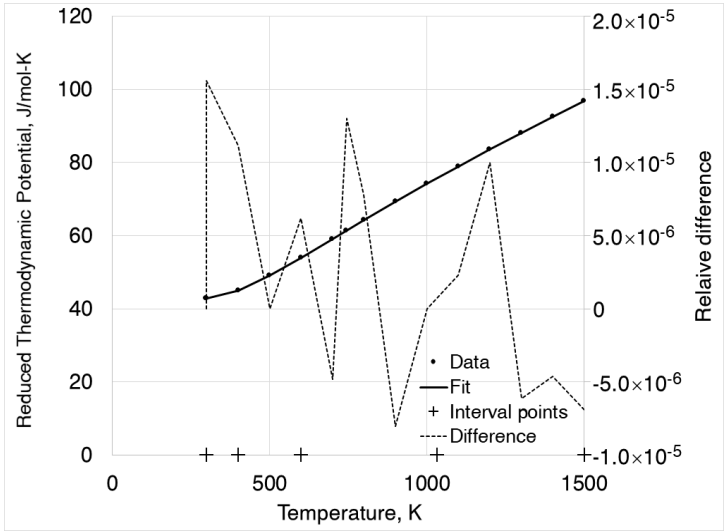
Lithium Hydroxide (LiOH) – H1Li1O1(cr)

MW = 23.948 g/mol – $\Delta_f H^o$(298.15 K) = -484.926 kJ/mol – S^o(298.15 K) = 42.821 J/mol-K – p^o = 0.1 MPa



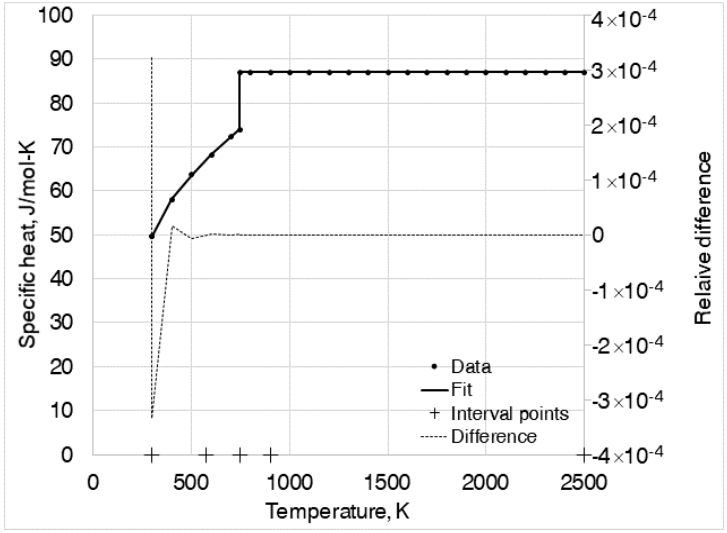


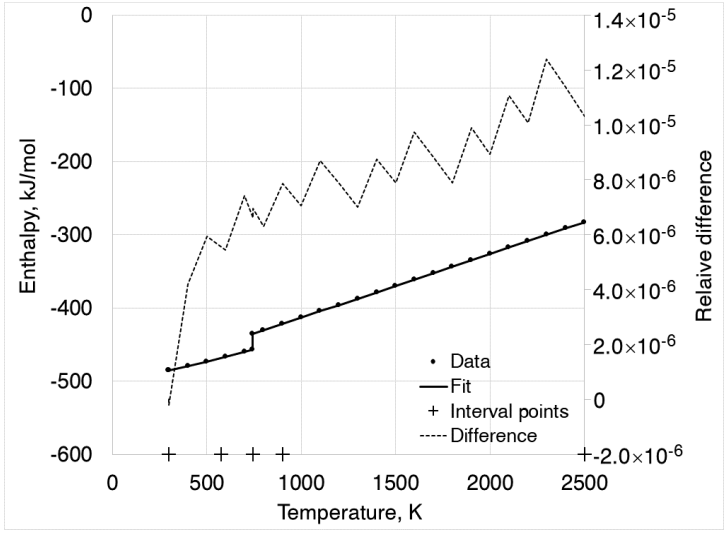


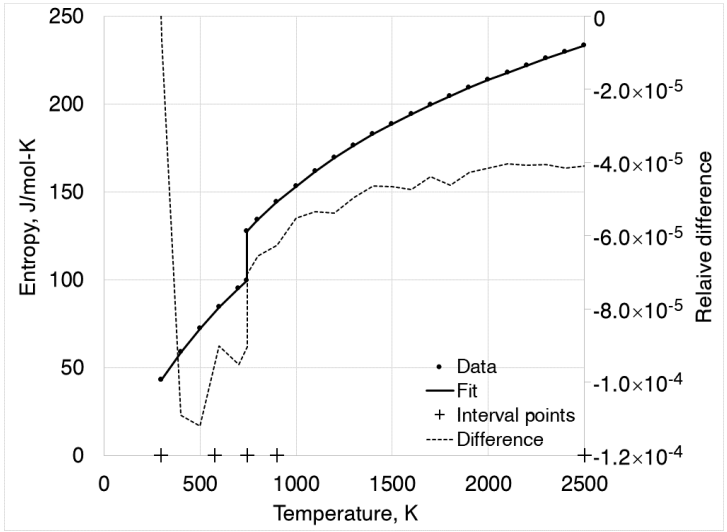


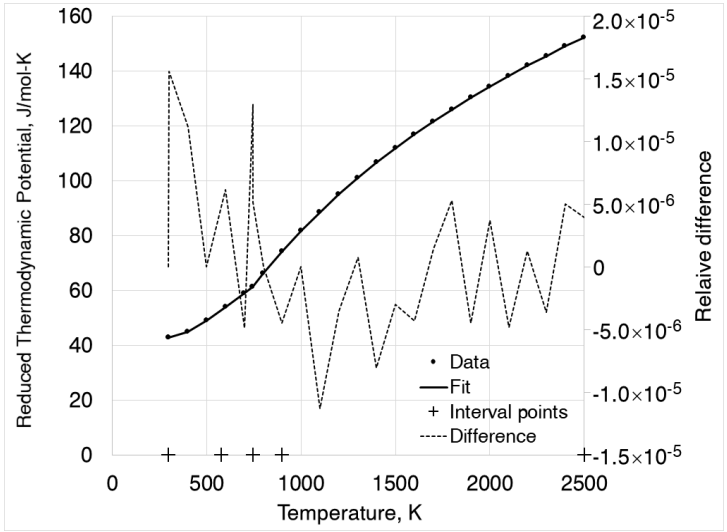
Lithium Hydroxide (LiOH) – H1Li1O1(cr,l)

MW = 23.948 g/mol – $\Delta_f H^o$(298.15 K) = -484.926 kJ/mol – S^o(298.15 K) = 42.821 J/mol-K – p^o = 0.1 MPa



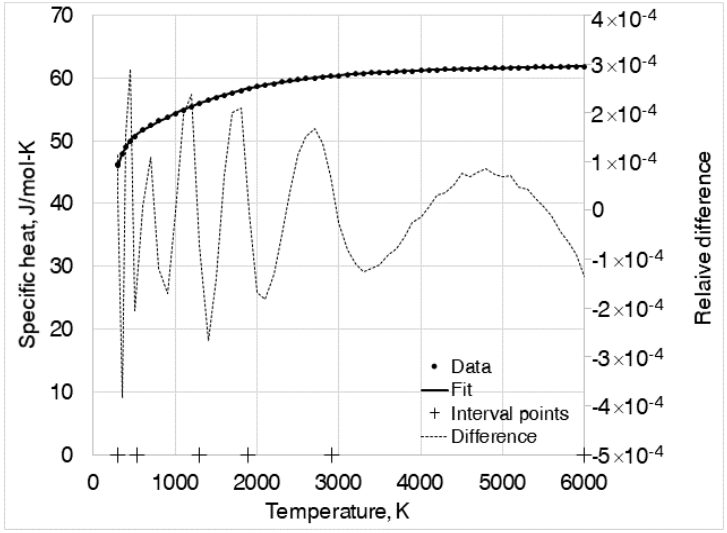


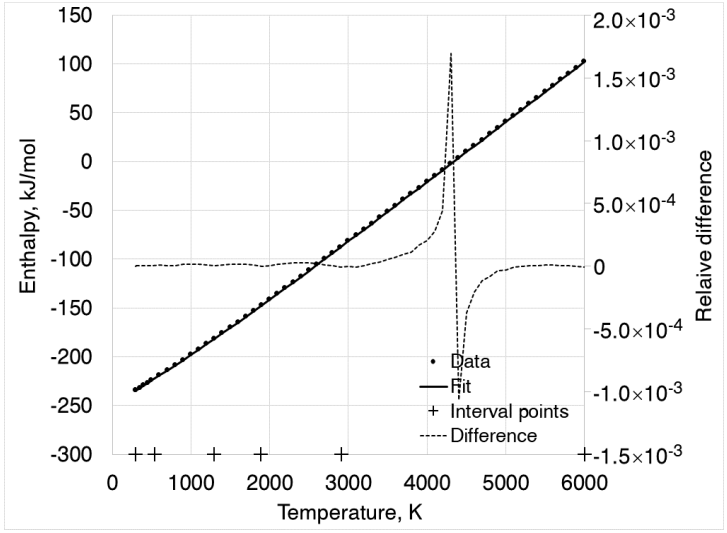


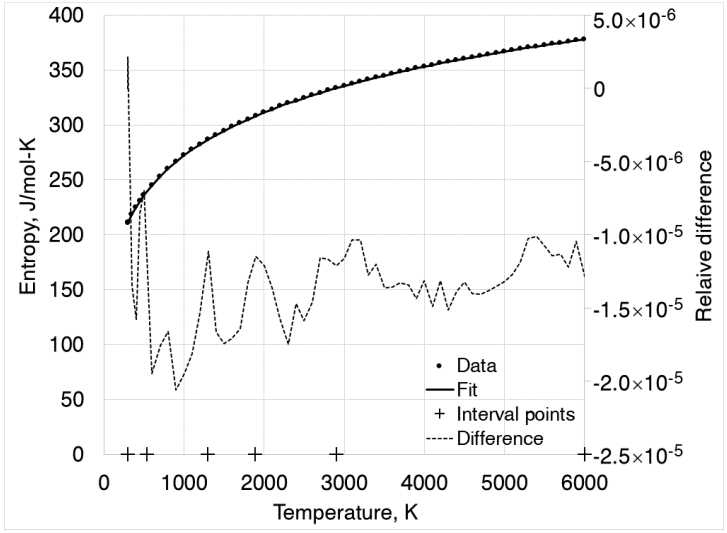


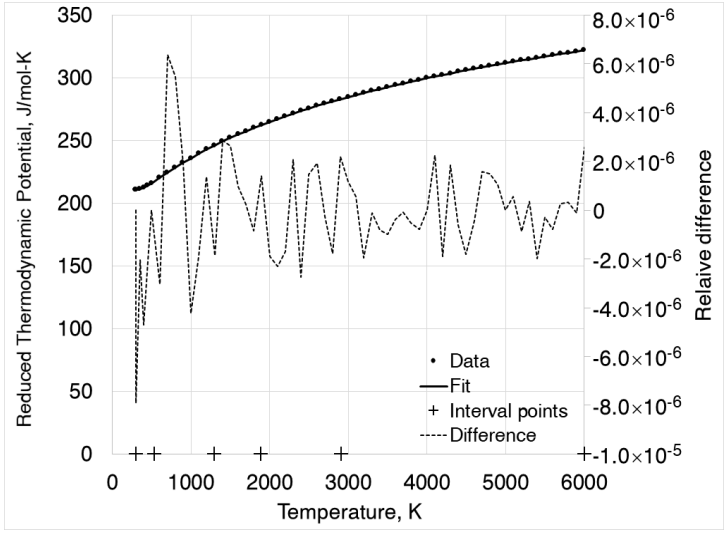
Lithium Hydroxide (LiOH) – H1Li1O1(g)

MW = 23.948 g/mol – $\Delta_f H^o$(298.15 K) = -234.304 kJ/mol – S^o(298.15 K) = 210.659 J/mol-K – p^o = 0.1 MPa



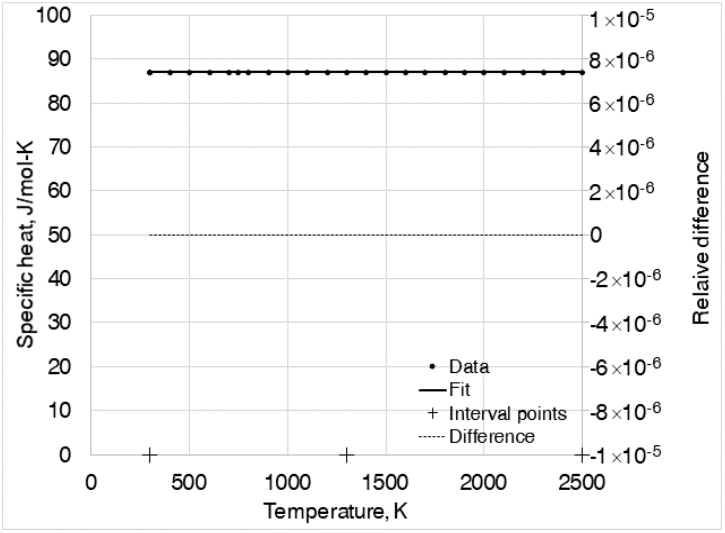


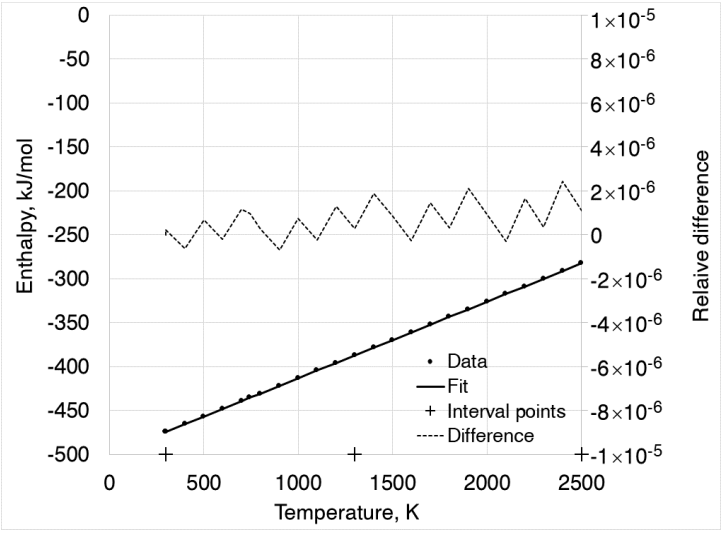


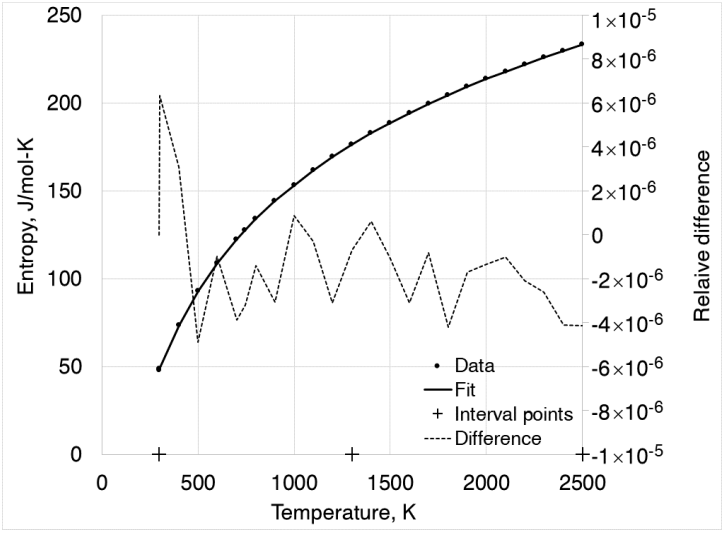


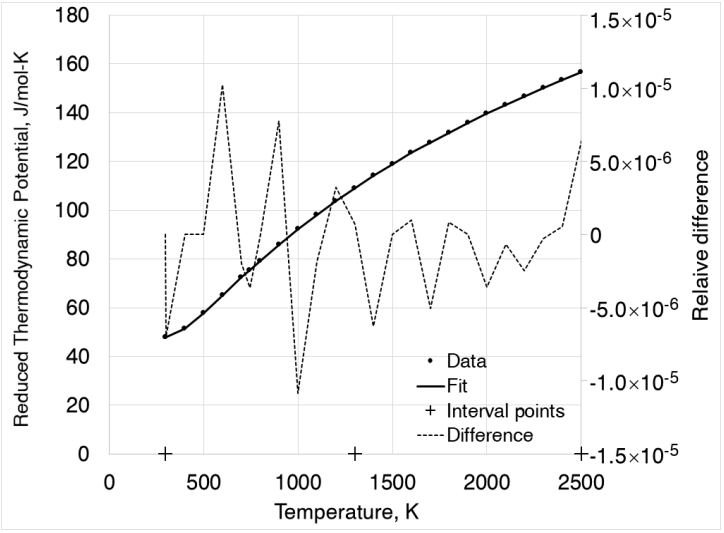
Lithium Hydroxide (LiOH) – H1Li1O1(l)

MW = 23.948 g/mol – $\Delta_f H^o$(298.15 K) = -474.422 kJ/mol – S^o(298.15 K) = 47.969 J/mol-K – p^o = 0.1 MPa



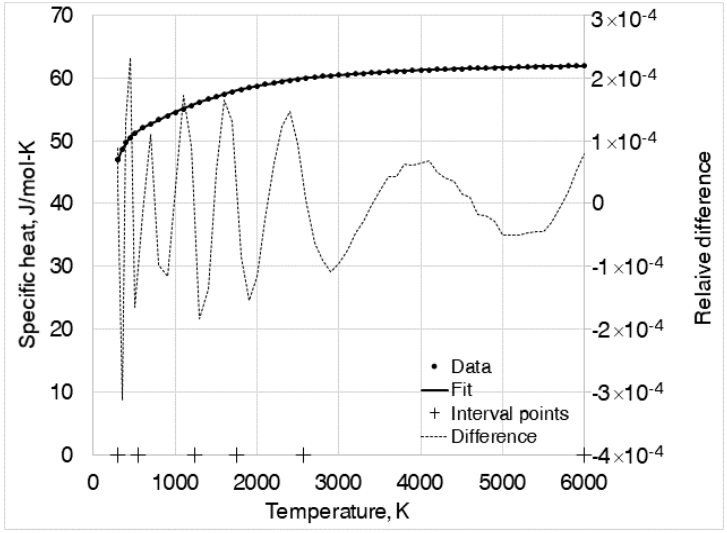


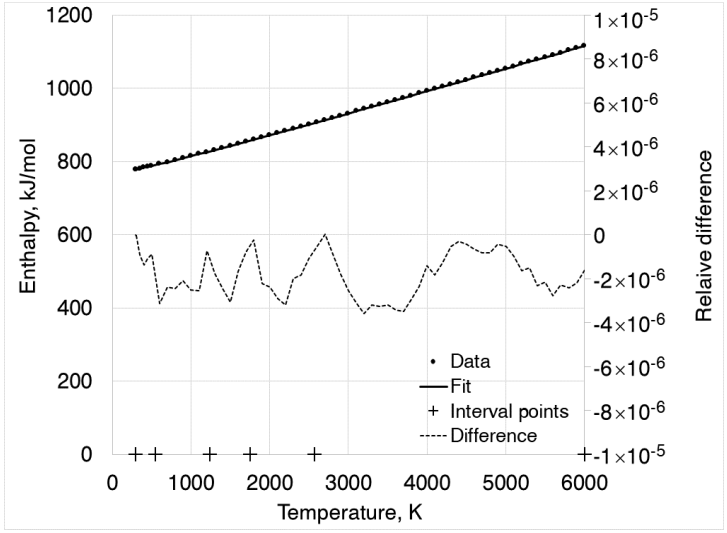


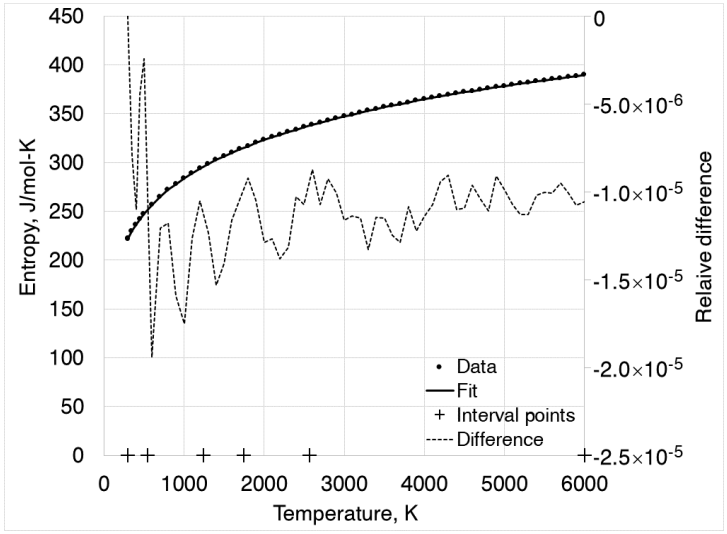


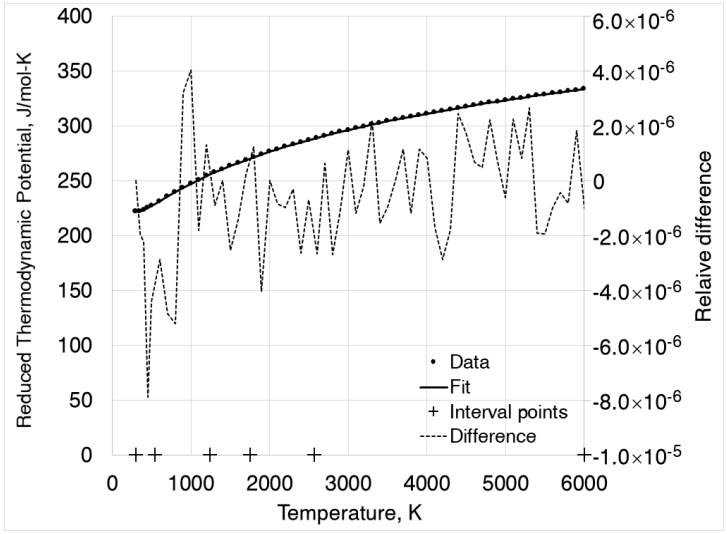
Lithium Hydroxide, Ion ($LiOH^+$) – H1Li1O1+(g)

MW = 23.948 g/mol – $\Delta_f H^o$(298.15 K) = 778.224 kJ/mol – S^o(298.15 K) = 221.685 J/mol-K – p^o = 0.1 MPa



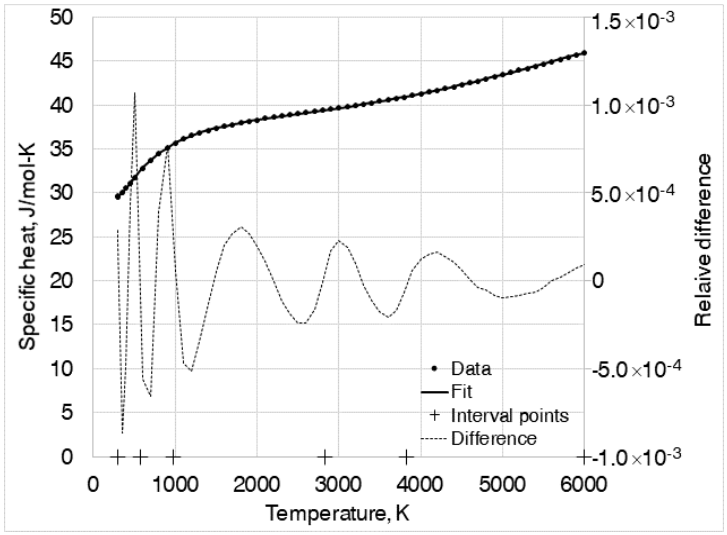


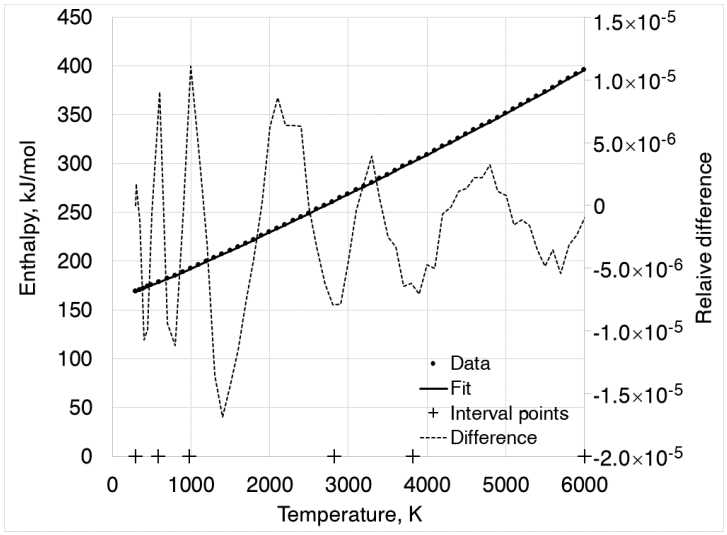


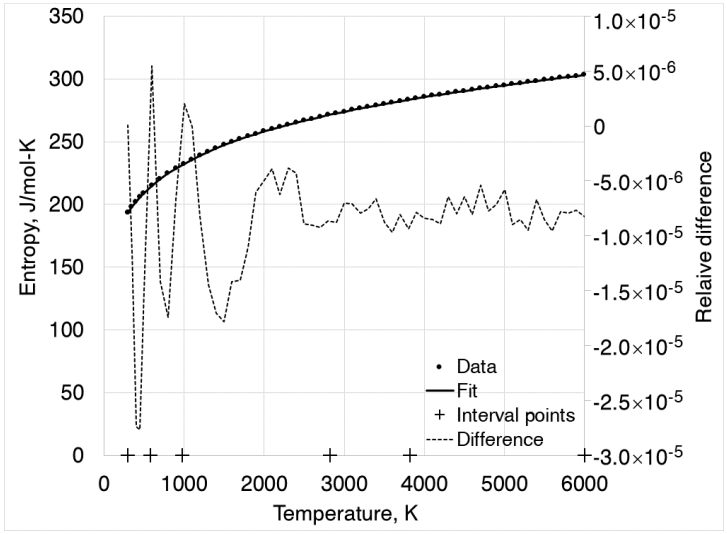


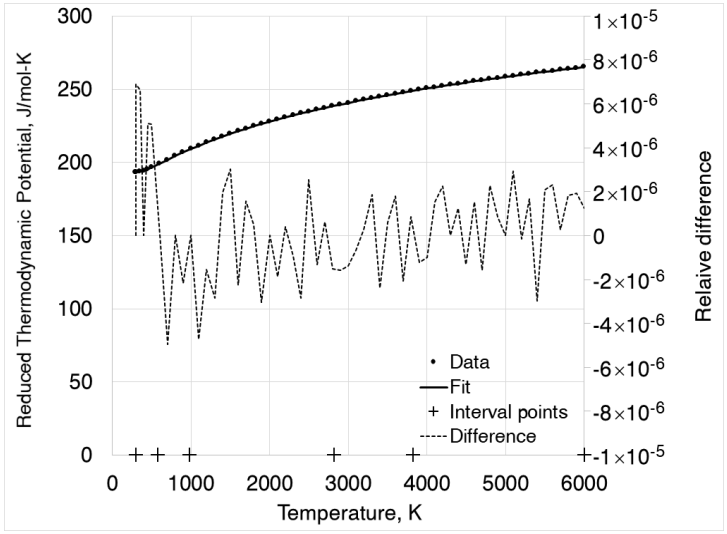
Magnesium Hydride (MgH) – H1Mg1(g)

MW = 25.313 g/mol – $\Delta_f H^o$(298.15 K) = 169.034 kJ/mol – S^o(298.15 K) = 193.2 J/mol-K – p^o = 0.1 MPa



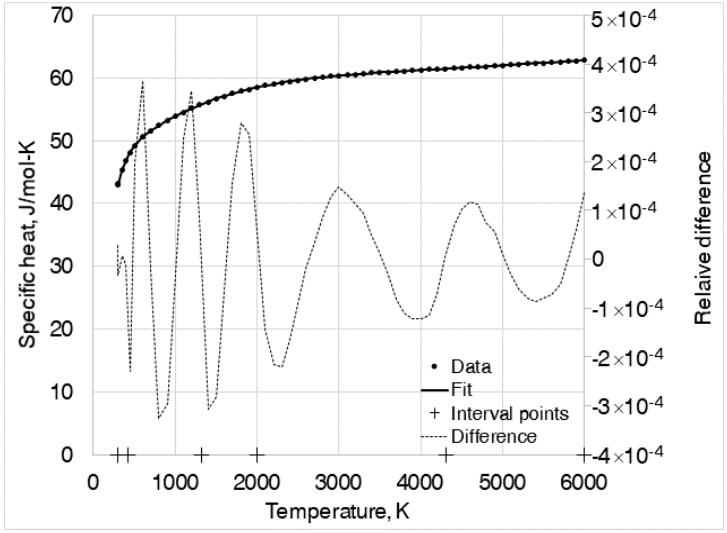


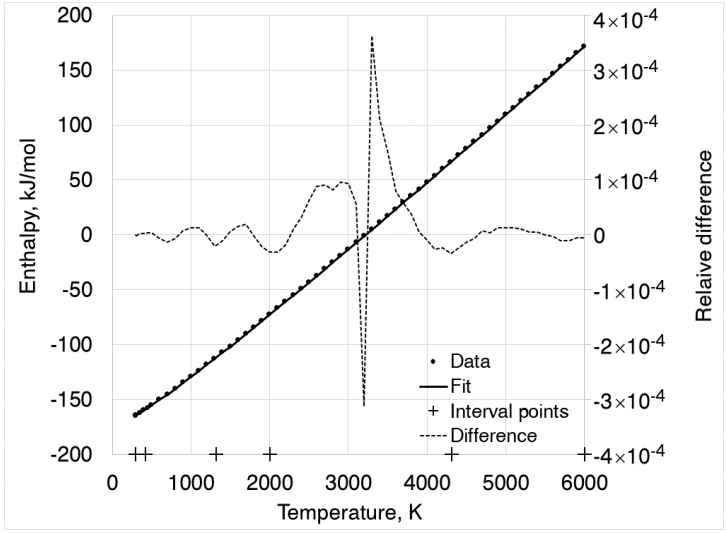


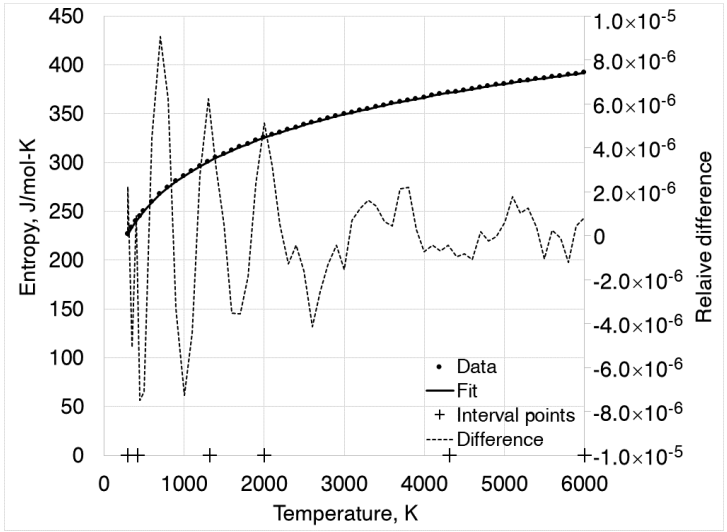


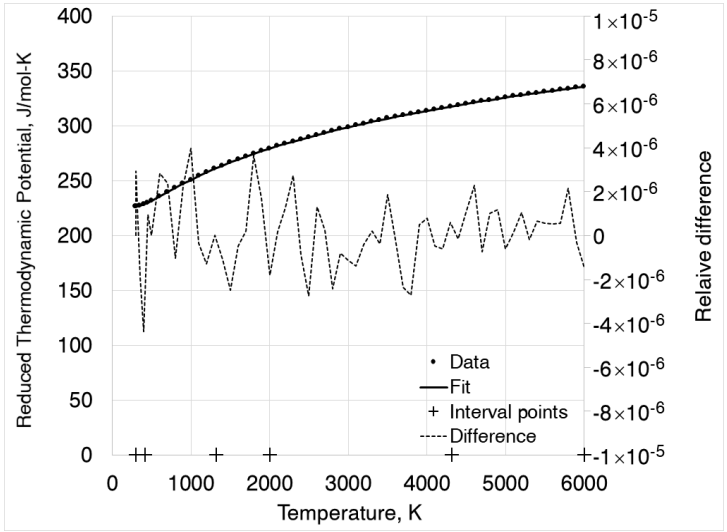
Magnesium Hydroxide (MgOH) – H1Mg1O1(g)

MW = 41.312 g/mol – $\Delta_f H^o$(298.15 K) = -164.755 kJ/mol – S^o(298.15 K) = 226.458 J/mol-K – p^o = 0.1 MPa



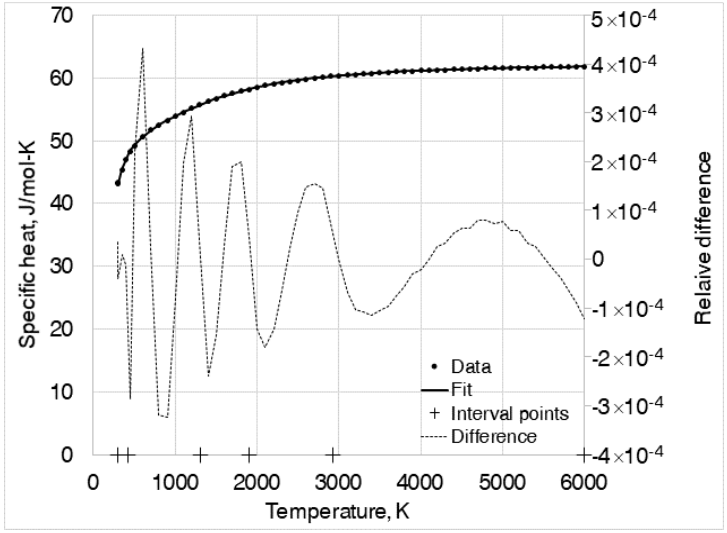


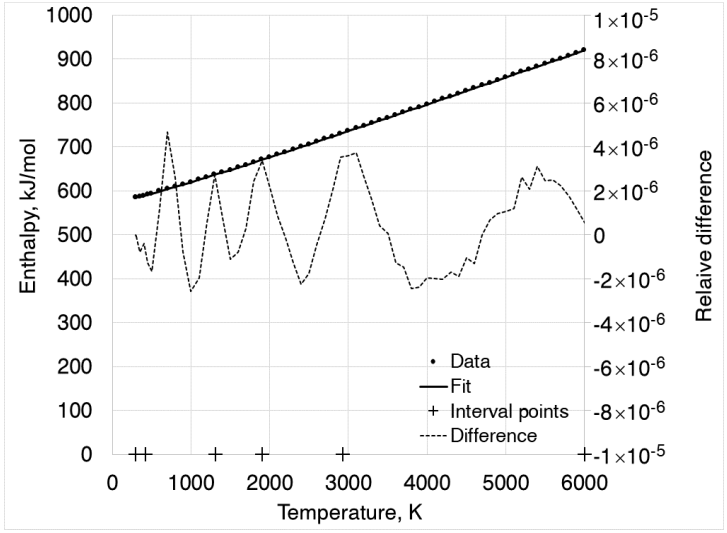


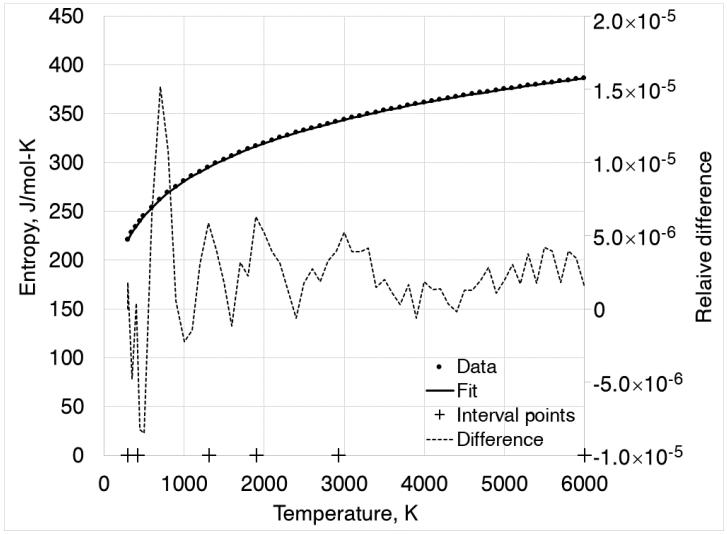


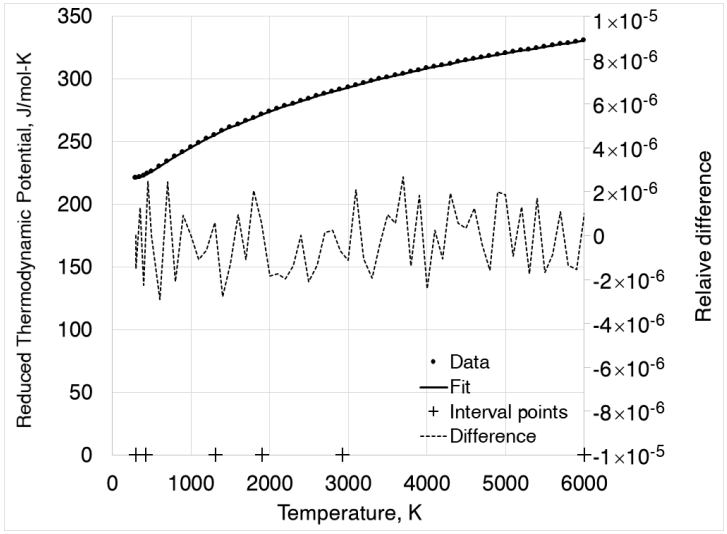
Magnesium Hydroxide, Ion ($MgOH^+$) – H1Mg1O1+(g)

MW = 41.312 g/mol – $\Delta_f H^o$(298.15 K) = 584.421 kJ/mol – S^o(298.15 K) = 220.824 J/mol-K – p^o = 0.1 MPa



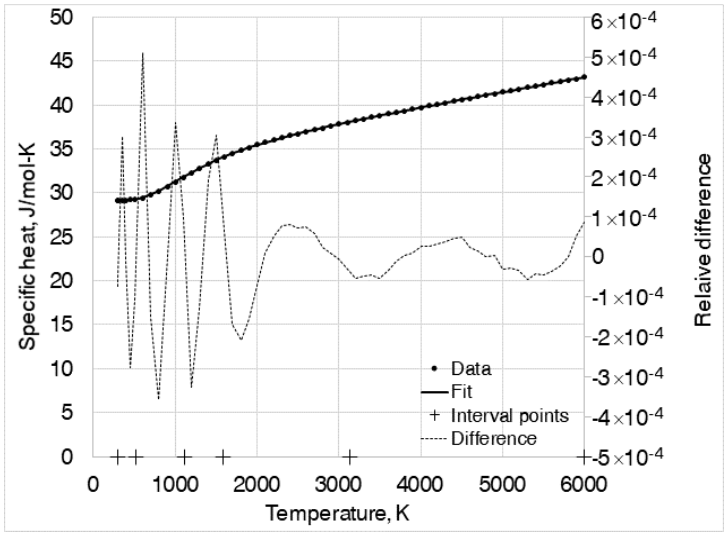


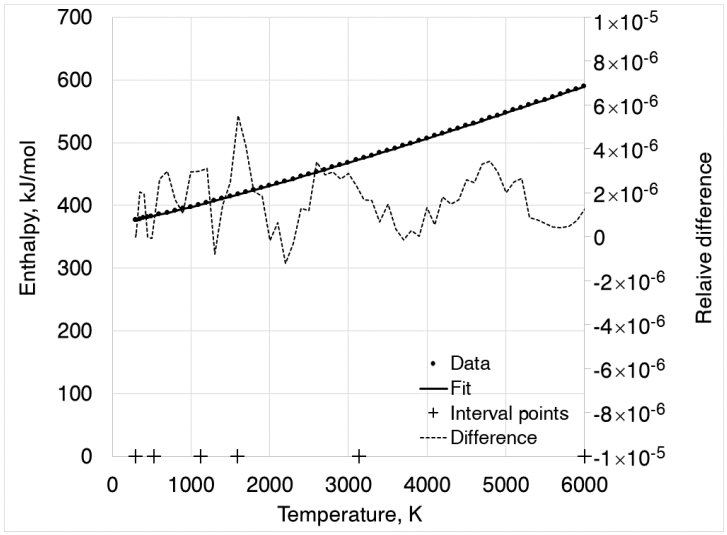


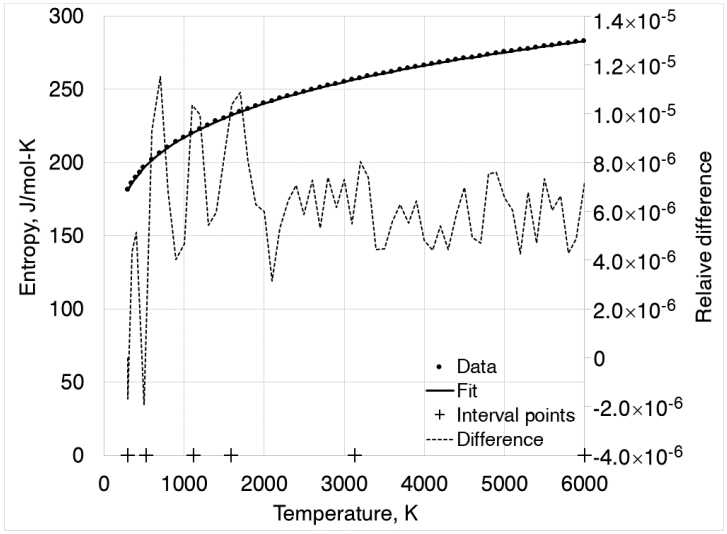


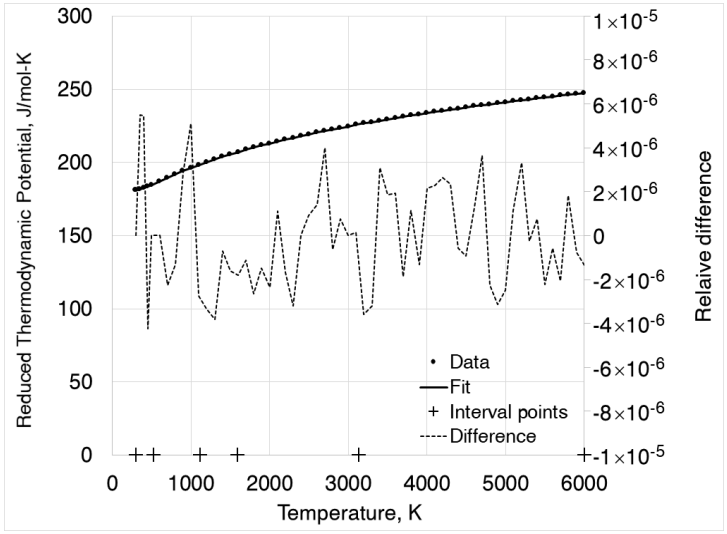
Imidogen (NH) – H1N1(g)

MW = 15.0146 g/mol – $\Delta_f H^o$(298.15 K) = 376.56 kJ/mol – S^o(298.15 K) = 181.253 J/mol-K – p^o = 0.1 MPa



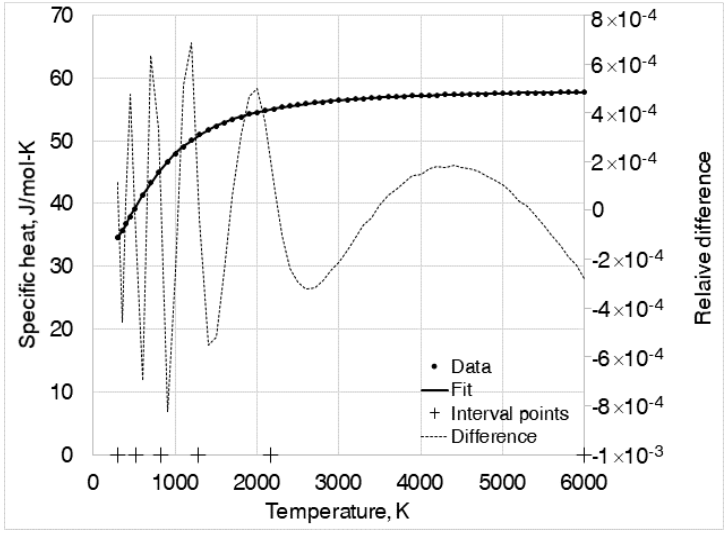


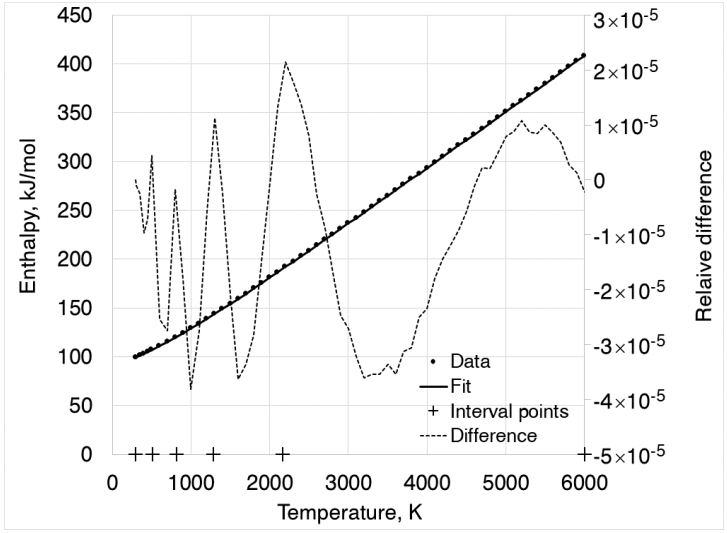


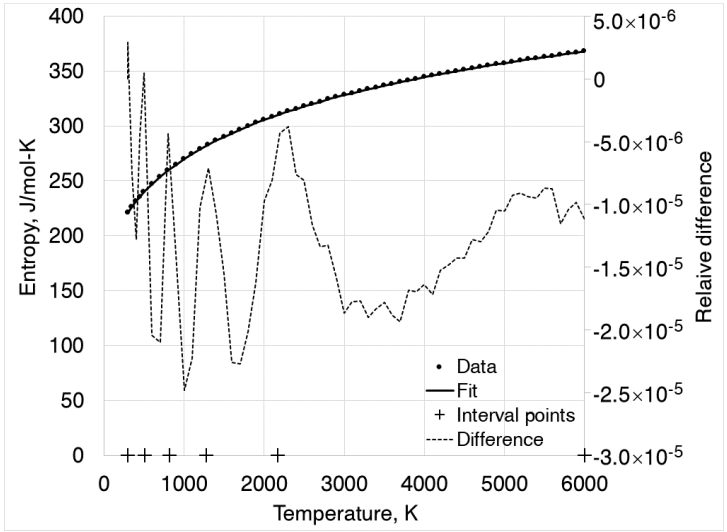


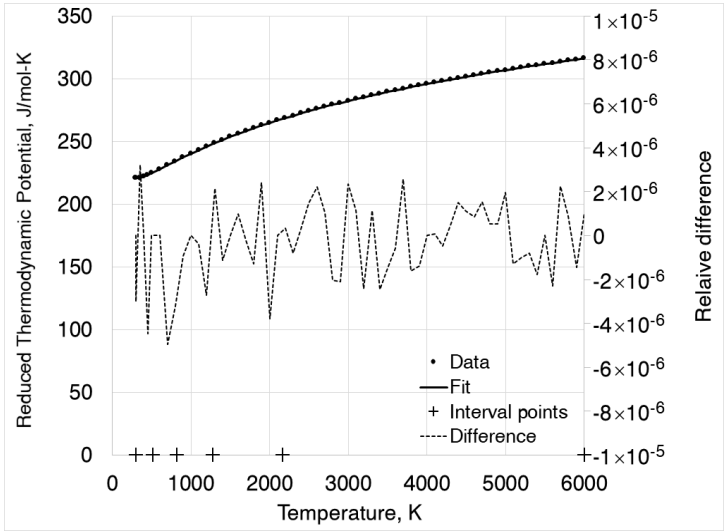
Nitrosyl Hydride (HNO) – H1N1O1(g)

MW = 31.014 g/mol – $\Delta_f H^o$(298.15 K) = 99.579 kJ/mol – S^o(298.15 K) = 220.722 J/mol-K – p^o = 0.1 MPa



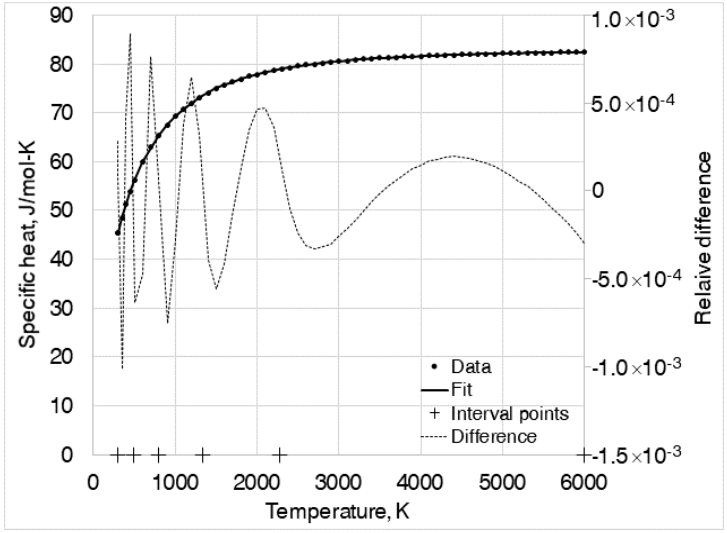


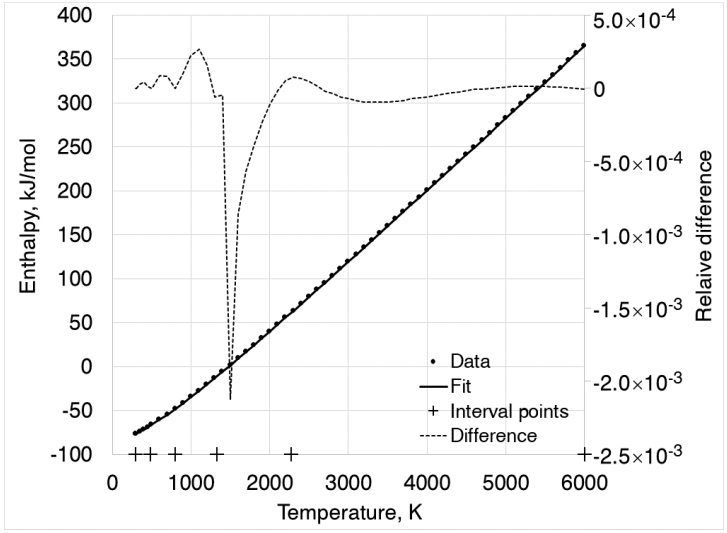


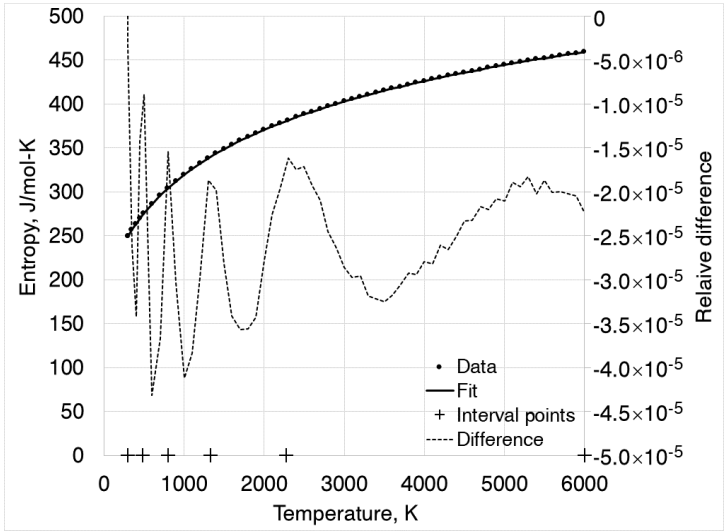


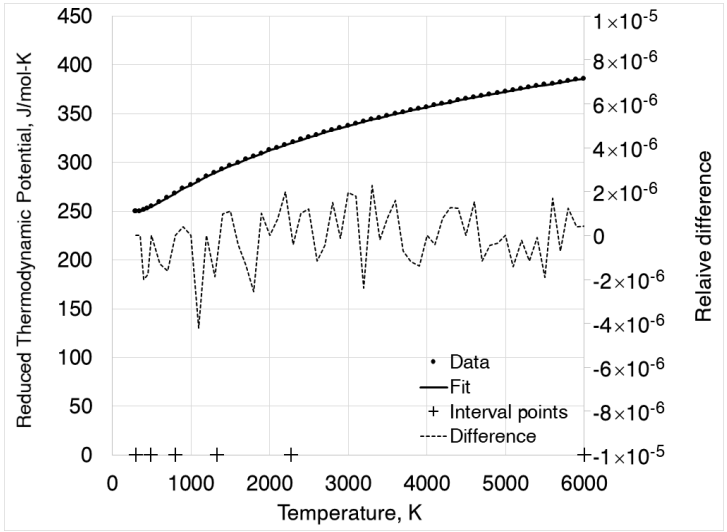
Nitrous Acid, Cis (HONO) – H1N1O2(g)_01

MW = 47.0134 g/mol – $\Delta_f H^o$(298.15 K) = -76.735 kJ/mol – S^o(298.15 K) = 249.416 J/mol-K – p^o = 0.1 MPa



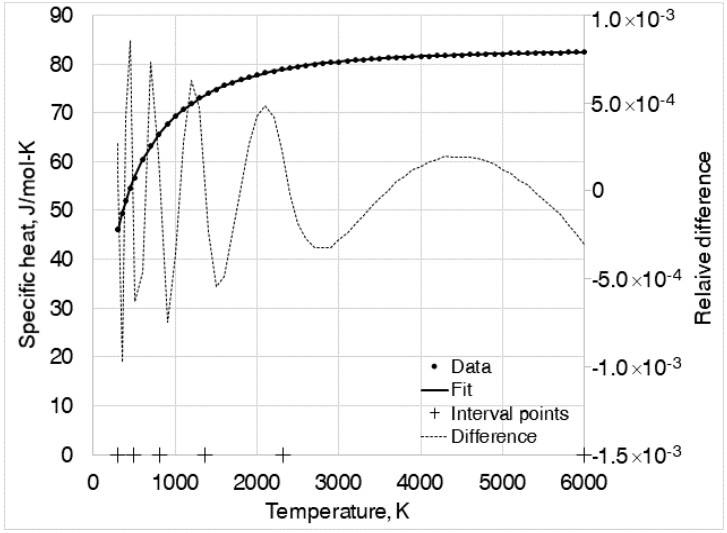


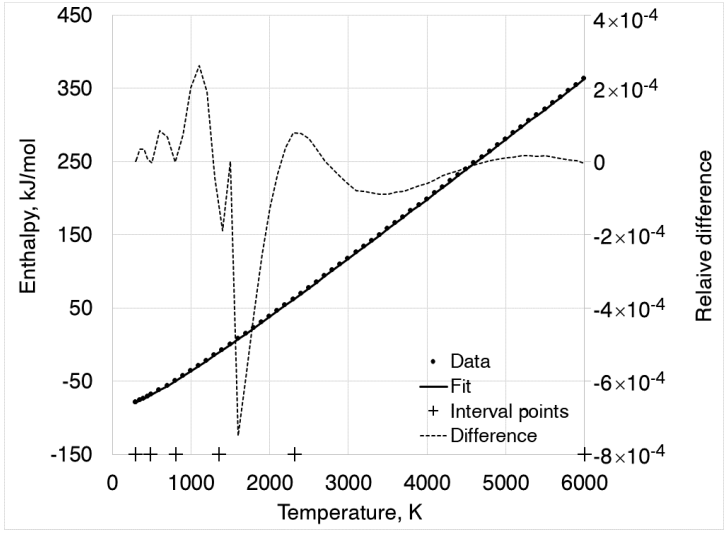


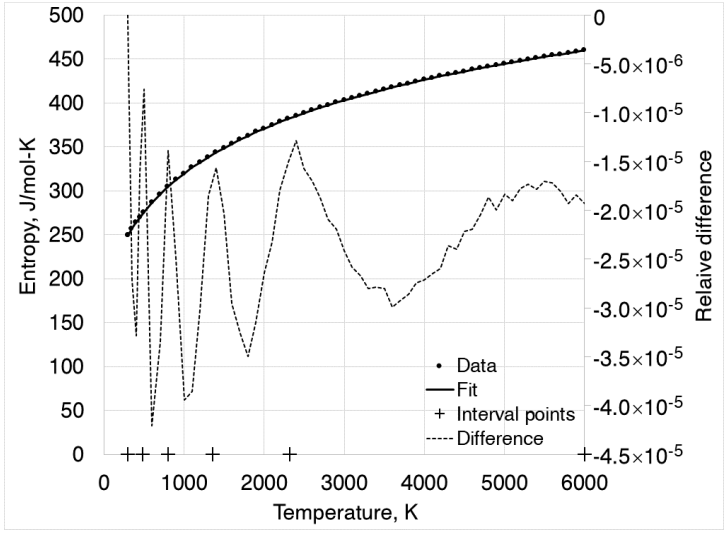


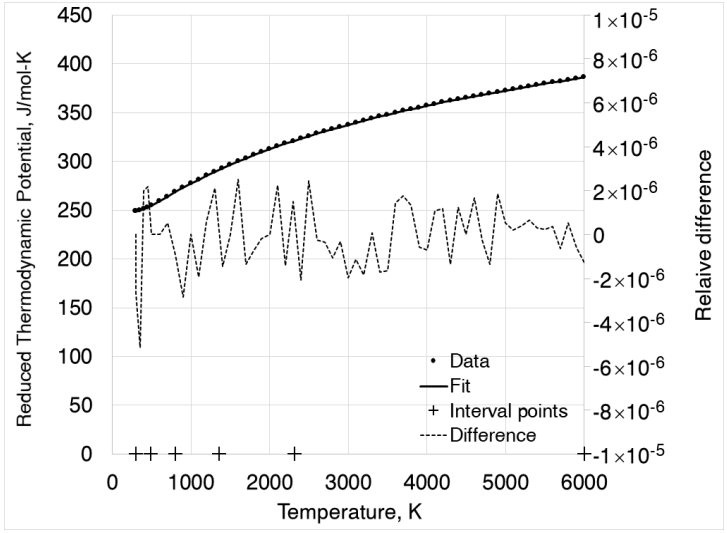
Nitrous Acid, Trans (HONO) – H1N1O2(g)_02

MW = 47.0134 g/mol – $\Delta_f H^o$(298.15 K) = -78.827 kJ/mol – S^o(298.15 K) = 249.261 J/mol-K – p^o = 0.1 MPa



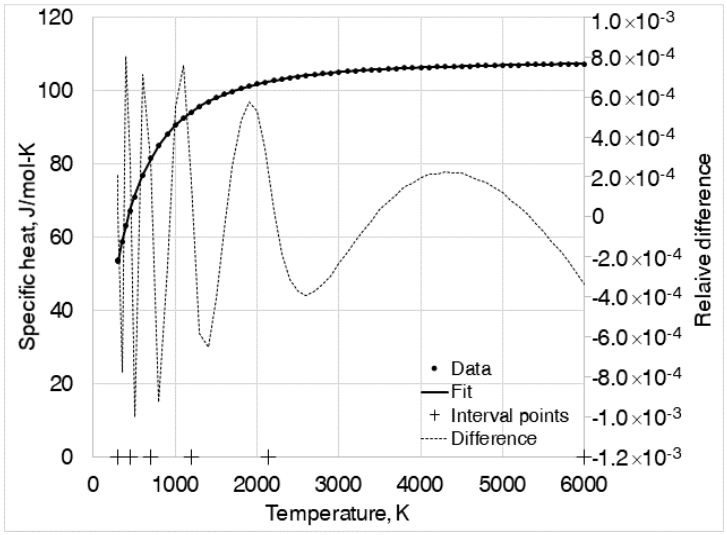


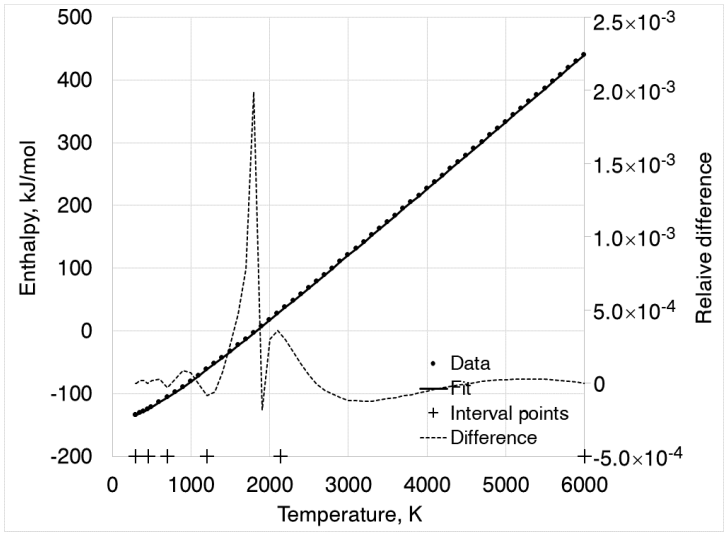


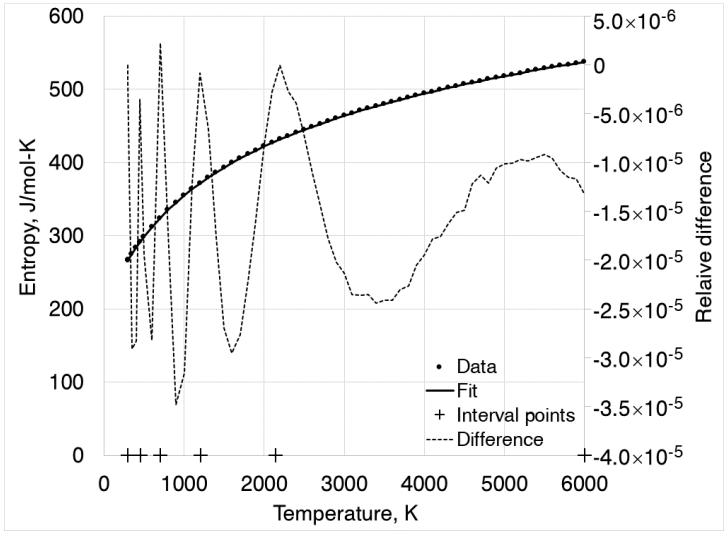


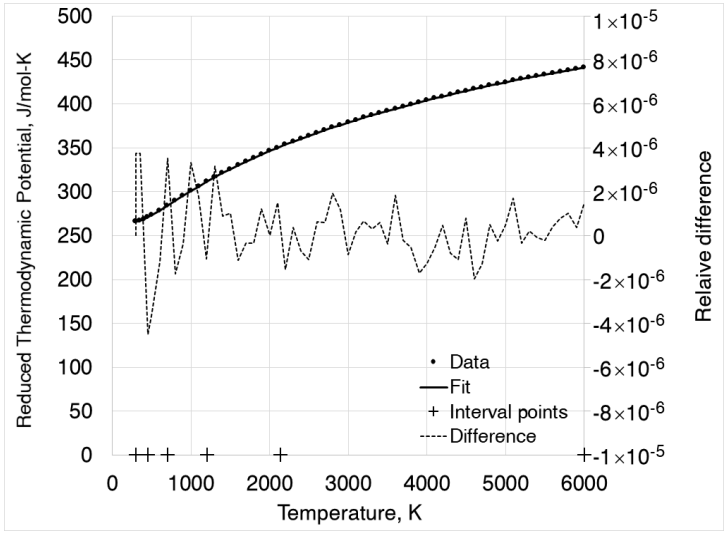
Nitric Acid ($HONO_2$) – H1N1O3(g)

MW = 63.0128 g/mol – $\Delta_f H^o$(298.15 K) = -134.306 kJ/mol – S^o(298.15 K) = 266.4 J/mol-K – p^o = 0.1 MPa



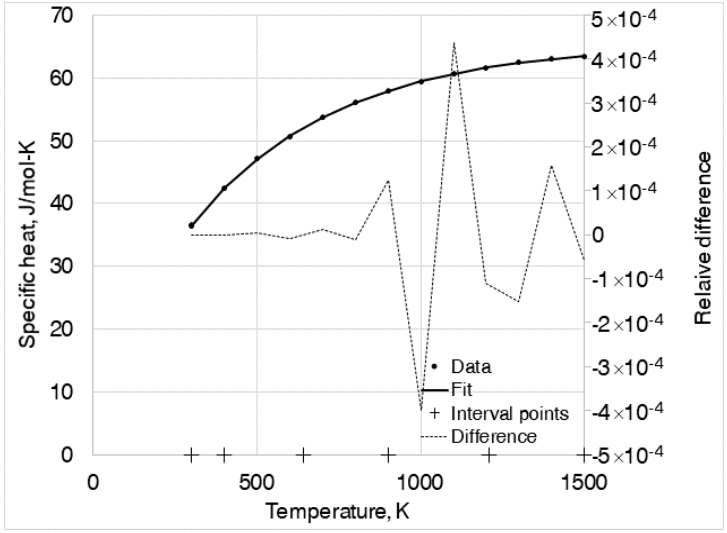


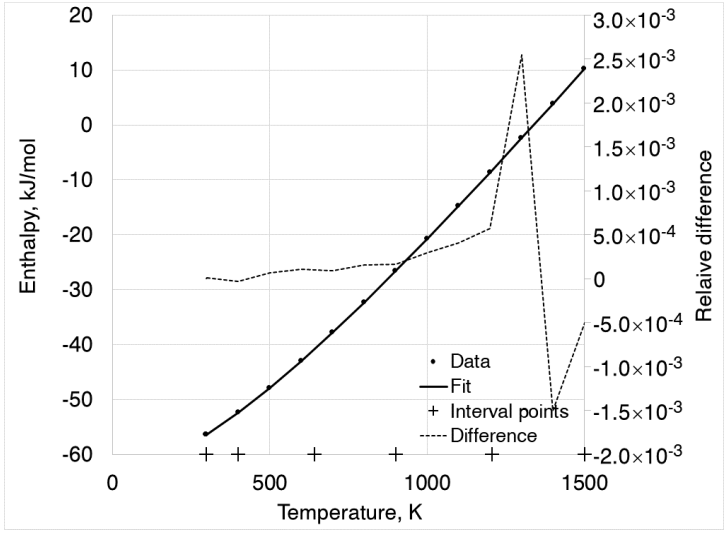


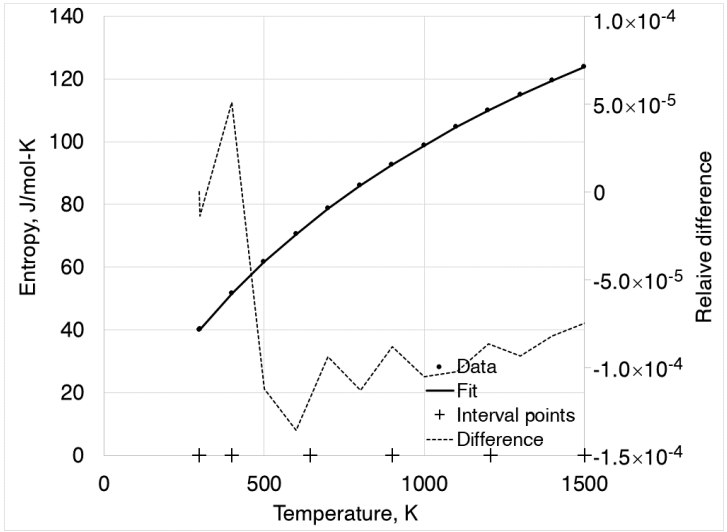


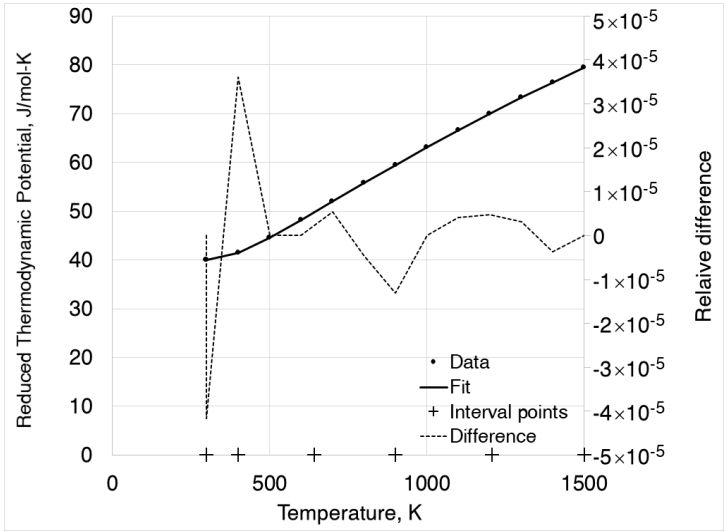
Sodium Hydride (NaH) – H1Na1(cr)

MW = 23.99771 g/mol – $\Delta_f H^o$(298.15 K) = -56.442 kJ/mol – S^o(298.15 K) = 40.016 J/mol-K – p^o = 0.1 MPa



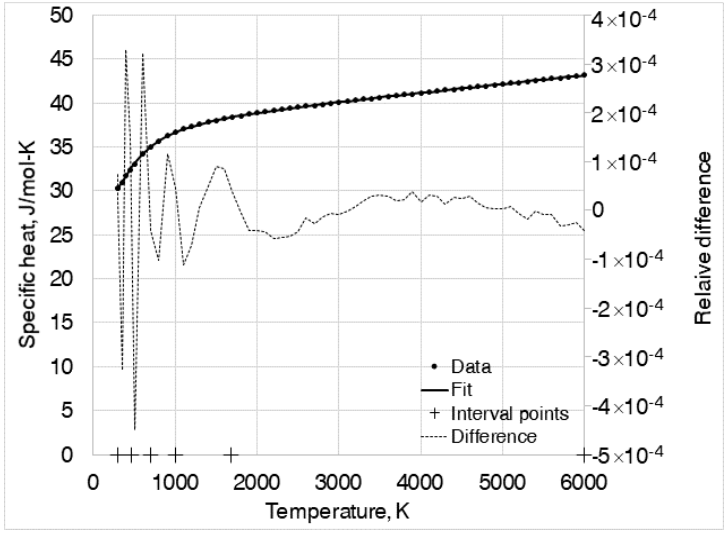


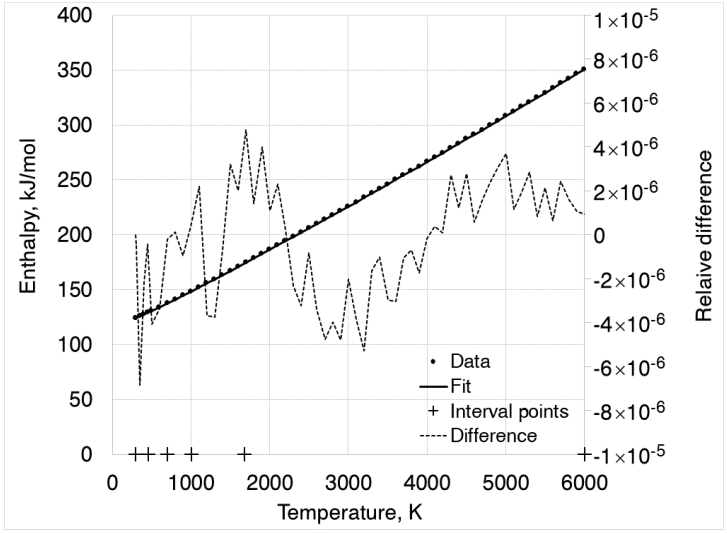


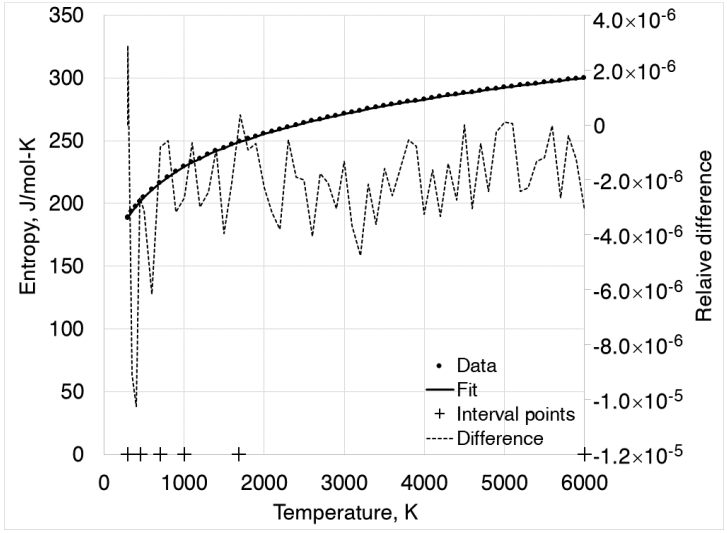


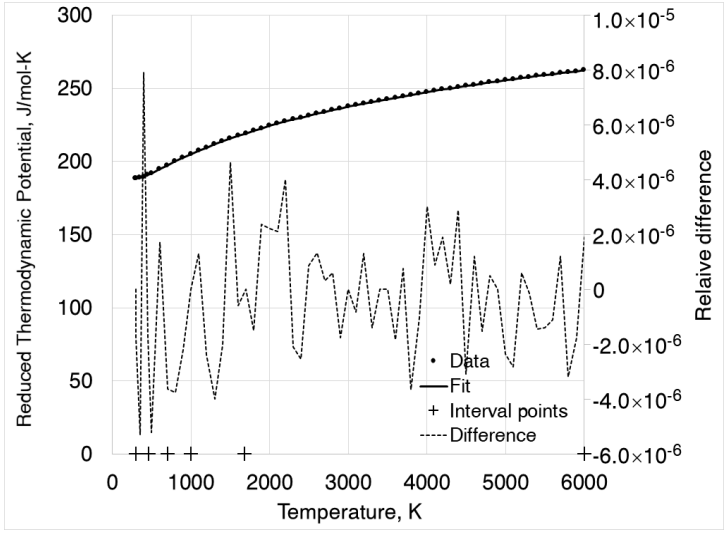
Sodium Hydride (NaH) – H1Na1(g)

MW = 23.99771 g/mol – $\Delta_f H^o$(298.15 K) = 124.265 kJ/mol – S^o(298.15 K) = 188.387 J/mol-K – p^o = 0.1 MPa



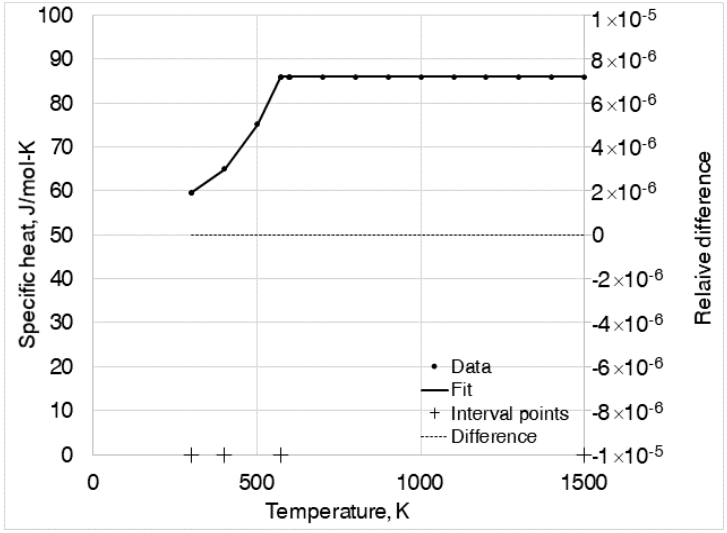


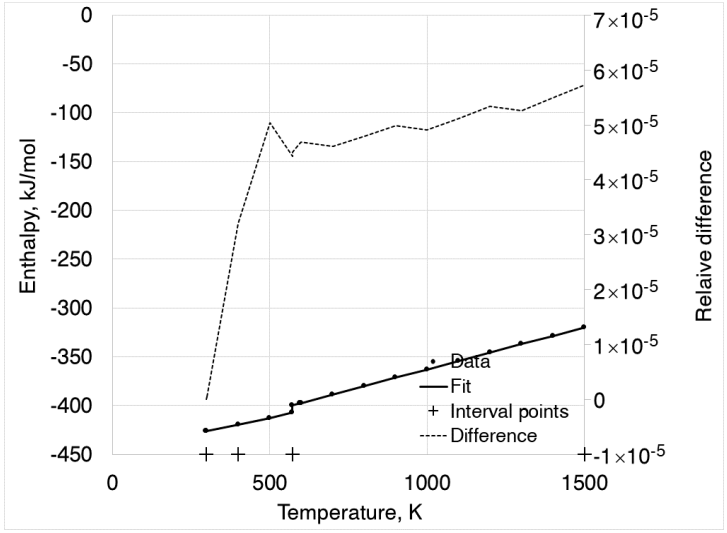


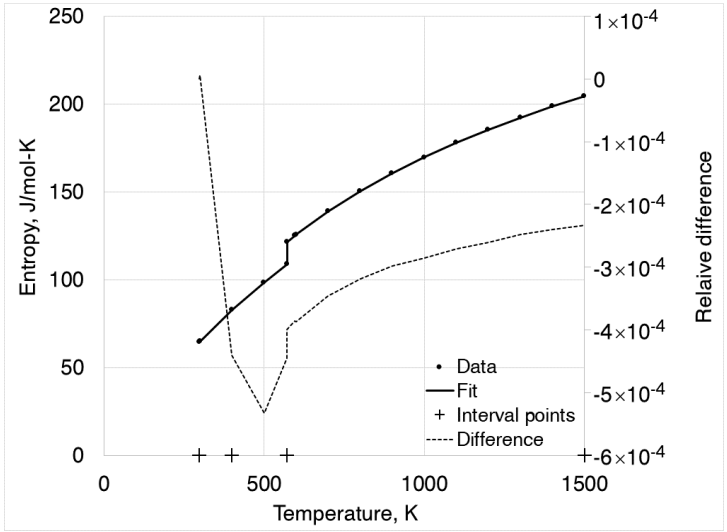


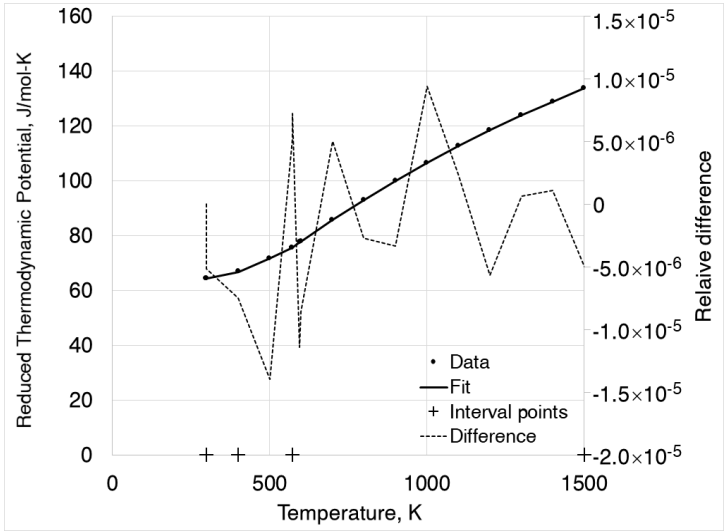
Sodium Hydroxide (NaOH) – H1Na1O1(cr)

MW = 39.9971 g/mol – $\Delta_f H^o$(298.15 K) = -425.931 kJ/mol – S^o(298.15 K) = 64.445 J/mol-K – p^o = 0.1 MPa



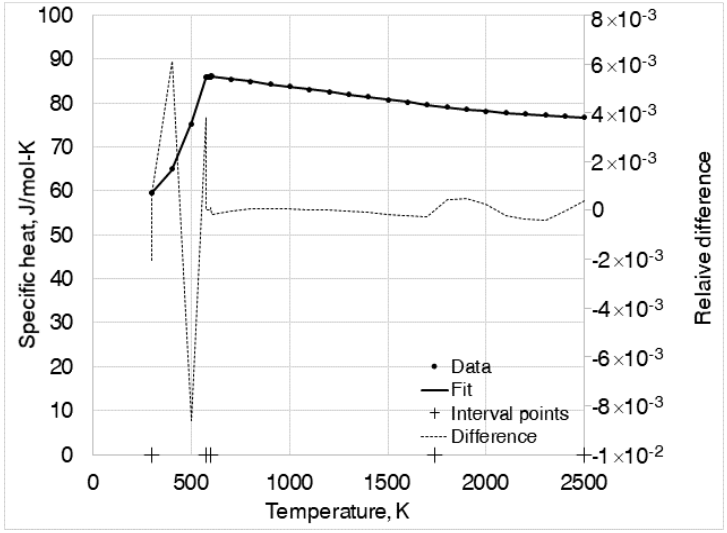


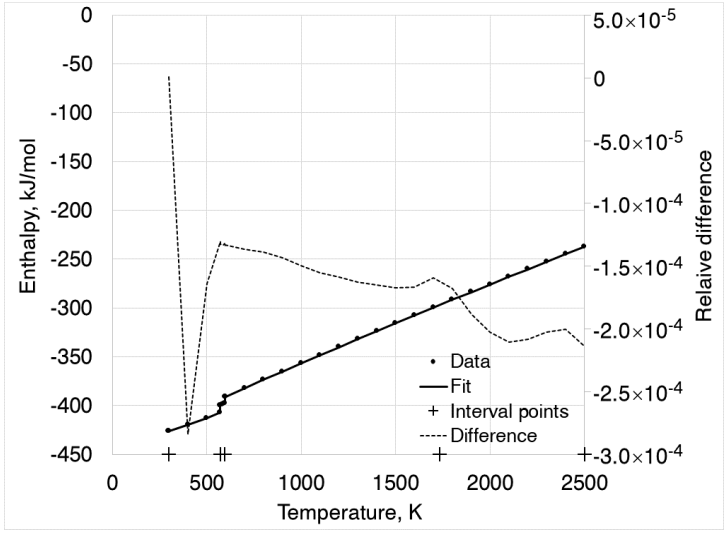


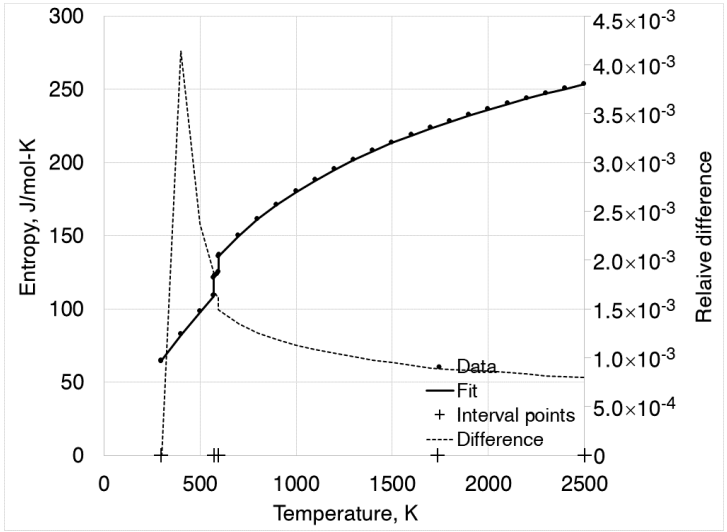


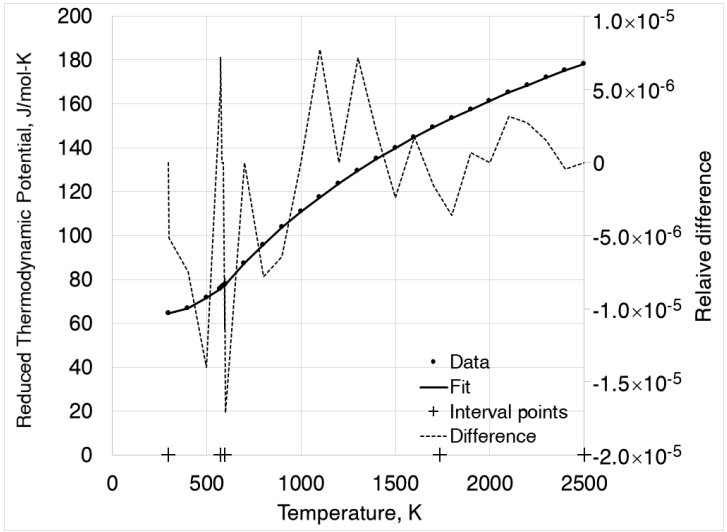
Sodium Hydroxide (NaOH) – H1Na1O1(cr,l)

MW = 39.9971 g/mol – $\Delta_f H^o$(298.15 K) = -425.931 kJ/mol – S^o(298.15 K) = 64.445 J/mol-K – p^o = 0.1 MPa



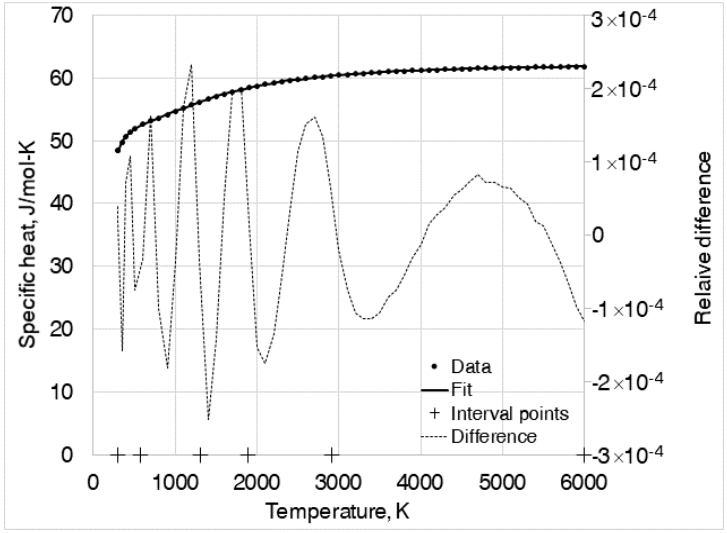


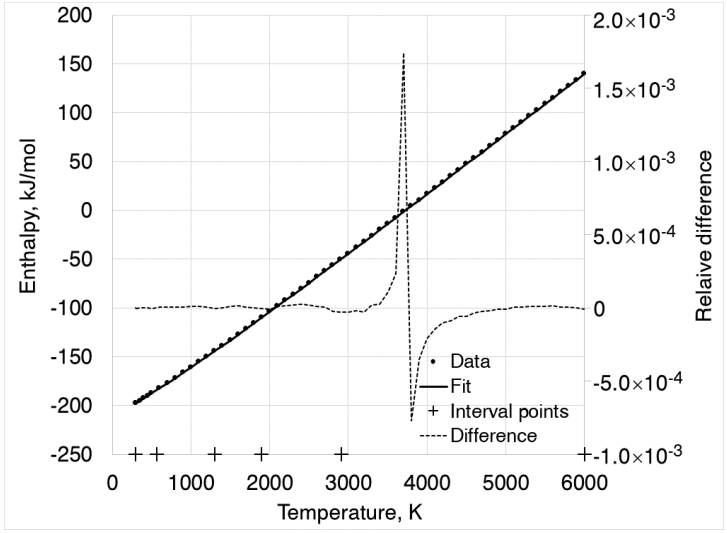


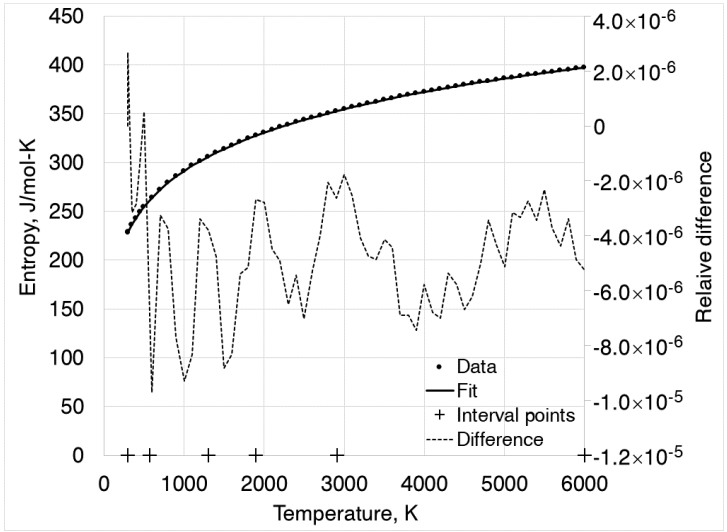


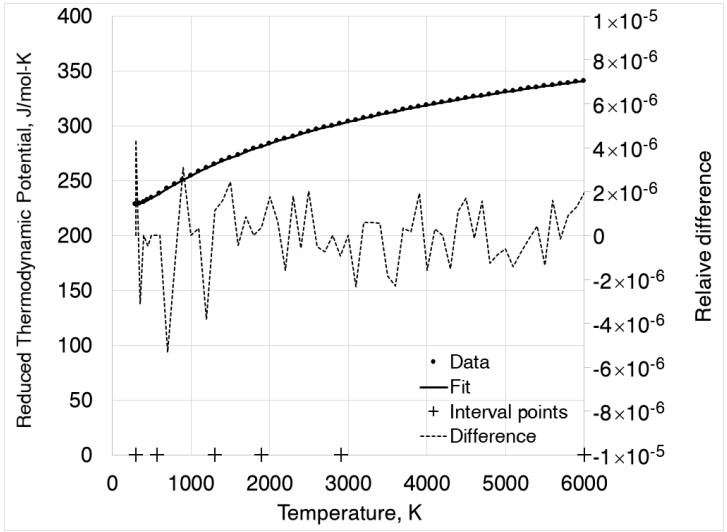
Sodium Hydroxide (NaOH) – H1Na1O1(g)

MW = 39.9971 g/mol – $\Delta_f H^o$(298.15 K) = -197.757 kJ/mol – S^o(298.15 K) = 228.443 J/mol-K – p^o = 0.1 MPa



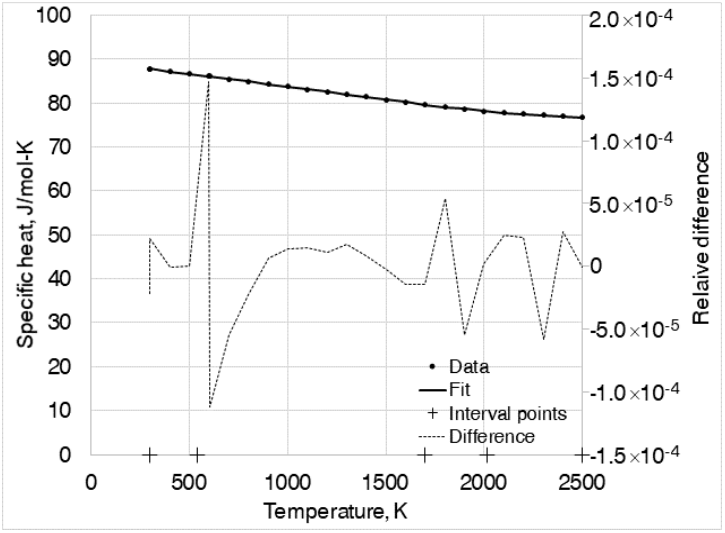


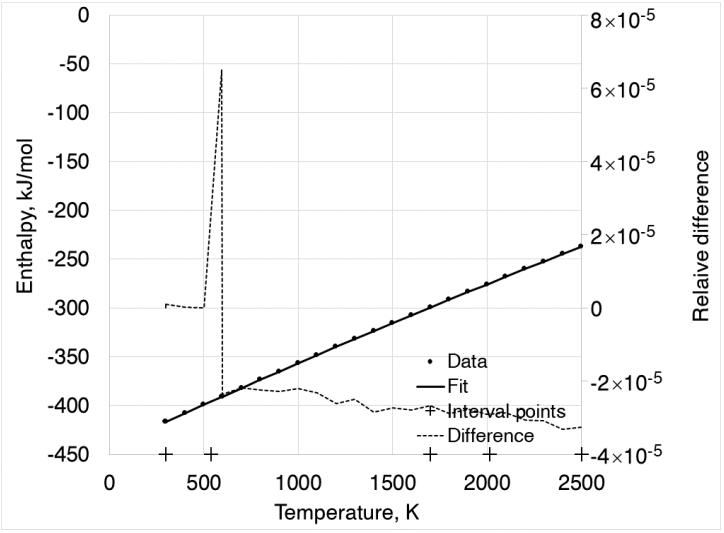


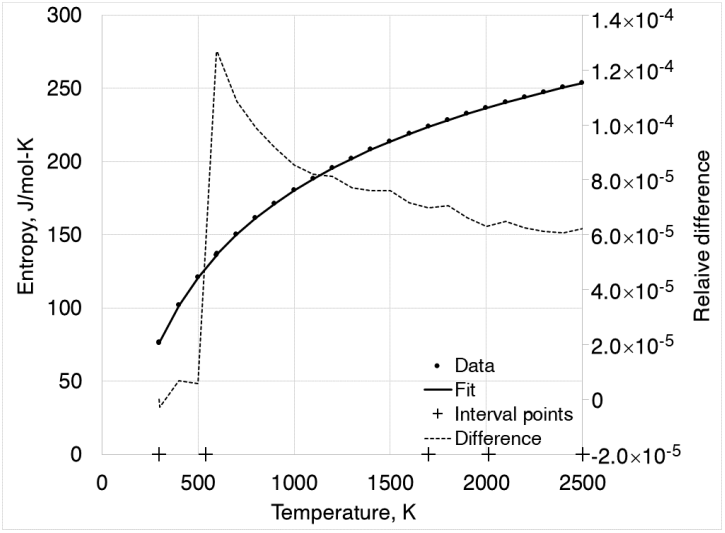


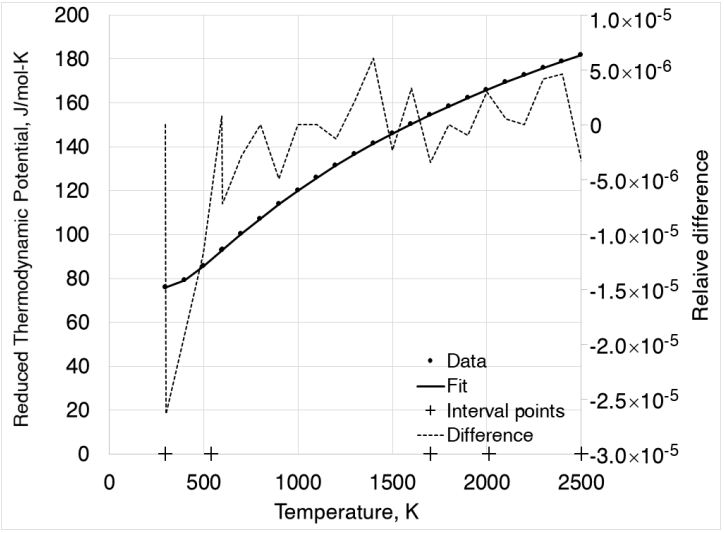
Sodium Hydroxide (NaOH) – H1Na1O1(l)

MW = 39.9971 g/mol – $\Delta_f H^o$(298.15 K) = -416.878 kJ/mol – S^o(298.15 K) = 75.891 J/mol-K – p^o = 0.1 MPa



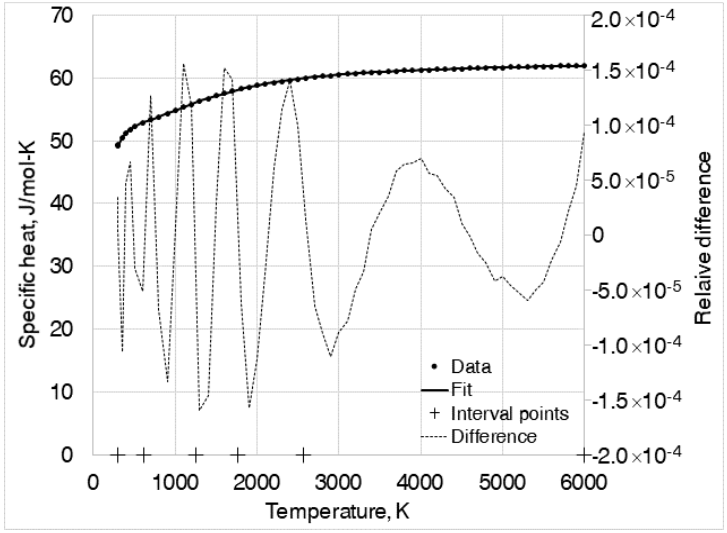


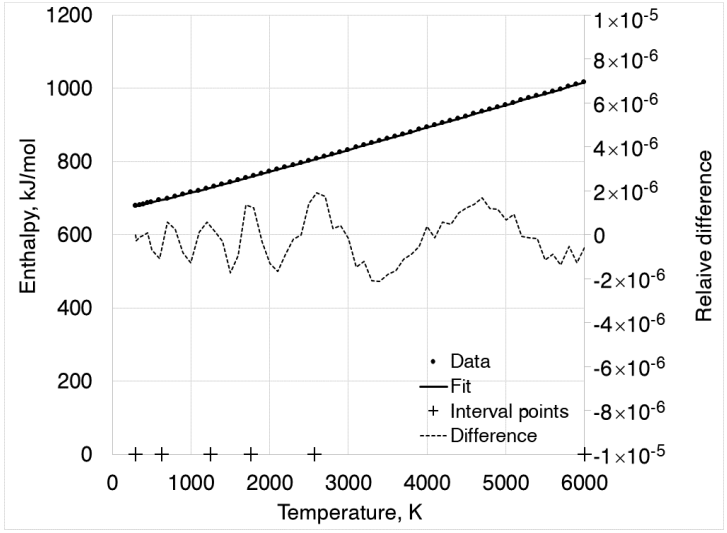


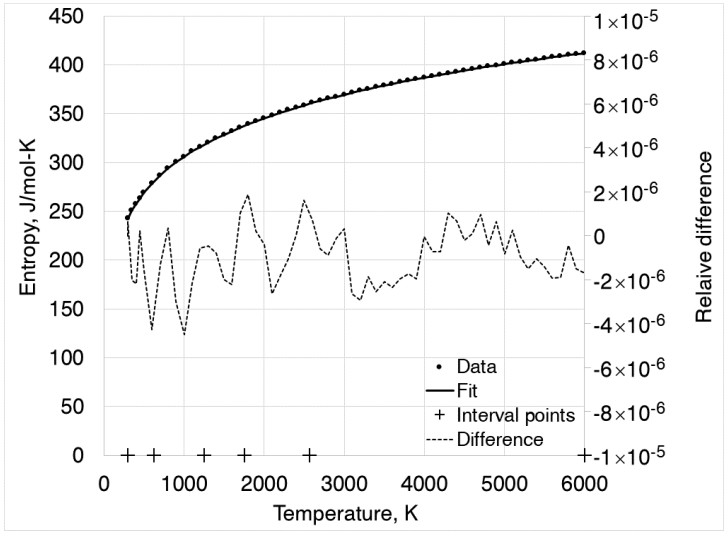


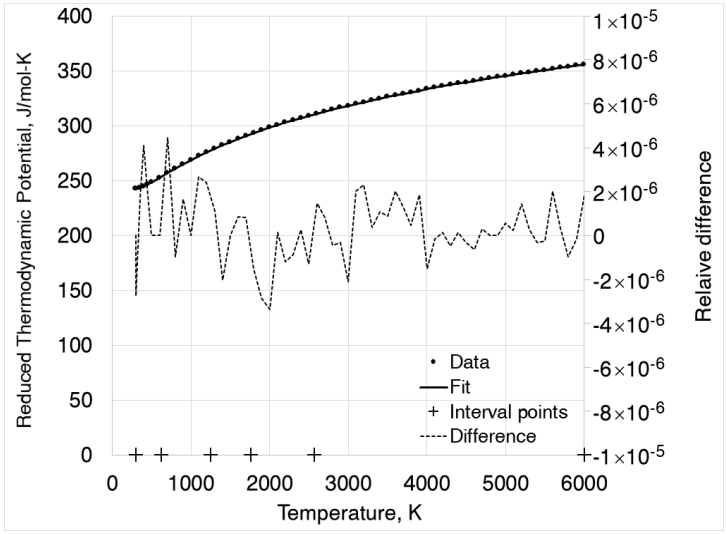
Sodium Hydroxide, Ion ($NaOH^+$) – H1Na1O1+(g)

MW = 39.9966 g/mol – $\Delta_f H^o$(298.15 K) = 677.808 kJ/mol – S^o(298.15 K) = 242.626 J/mol-K – p^o = 0.1 MPa



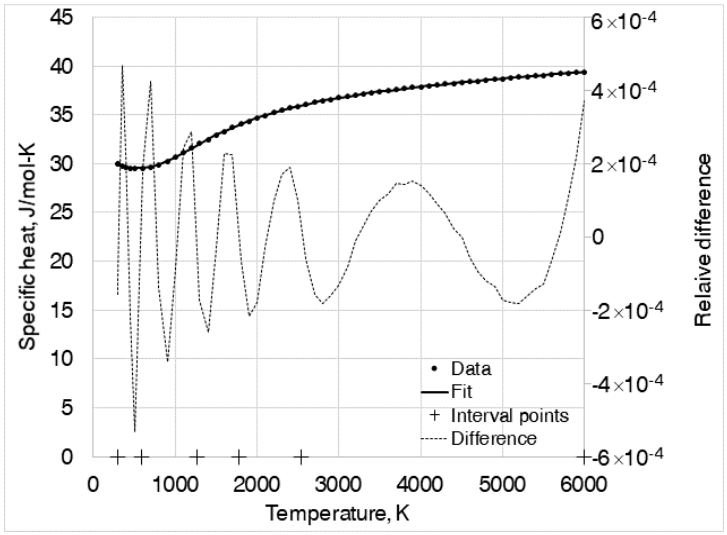


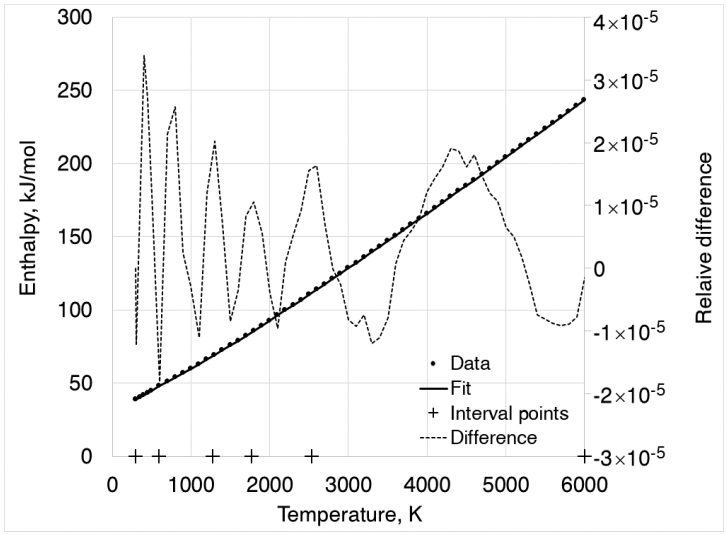


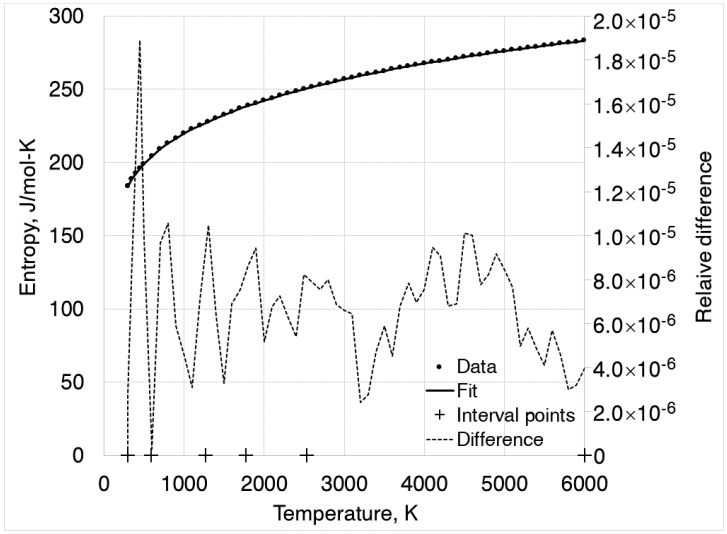


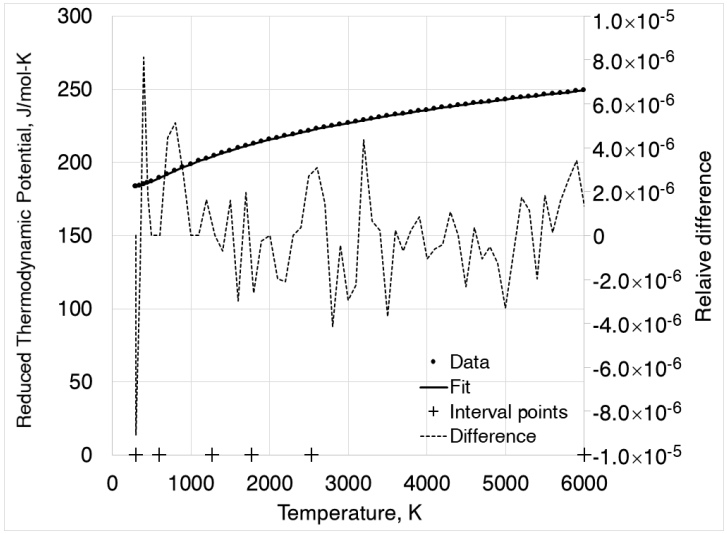
Hydroxyl (OH) – H1O1(g)

MW = 17.0073 g/mol – $\Delta_f H^o$(298.15 K) = 38.987 kJ/mol – S^o(298.15 K) = 183.708 J/mol-K – p^o = 0.1 MPa



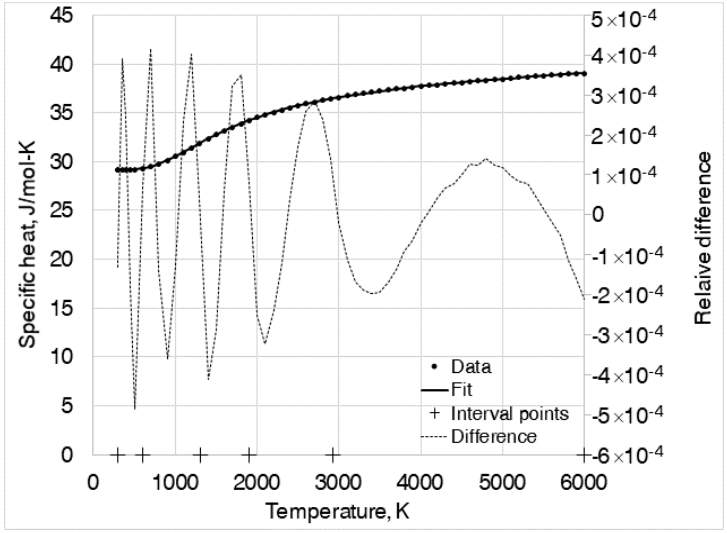


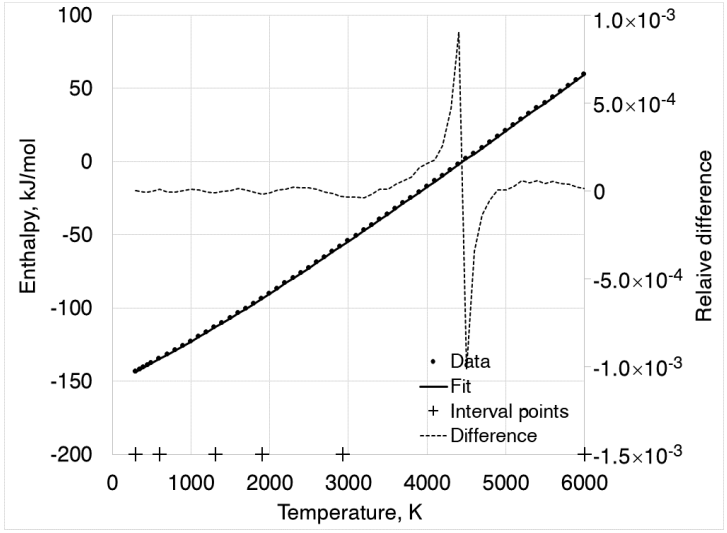


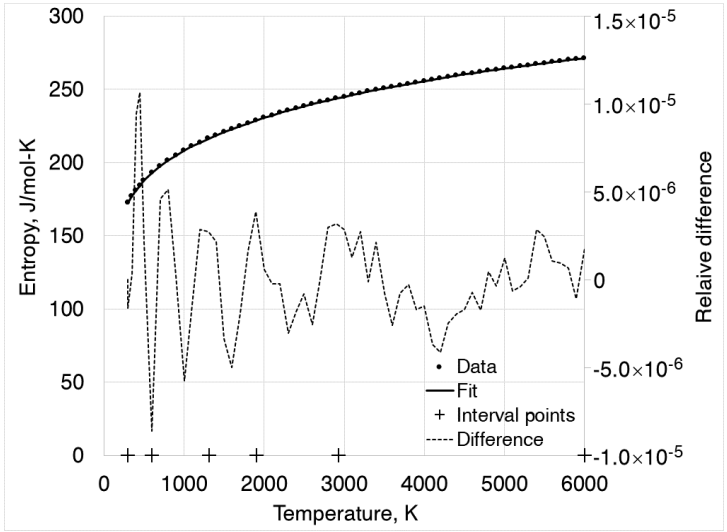


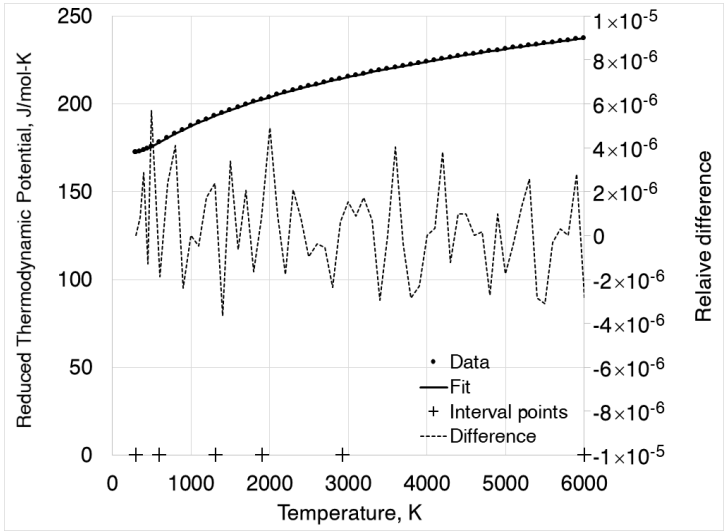
Hydroxyl, Ion (OH^-) – H1O1-(g)

MW = 17.0079 g/mol – $\Delta_f H^o$(298.15 K) = -143.595 kJ/mol – S^o(298.15 K) = 172.469 J/mol-K – p^o = 0.1 MPa



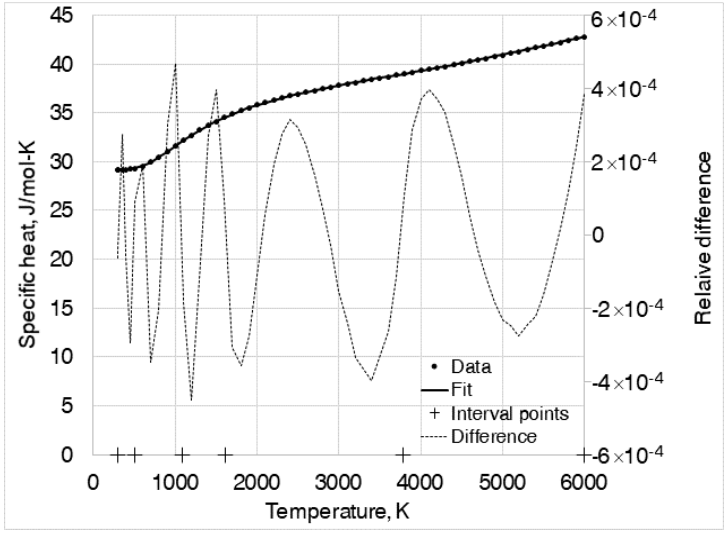


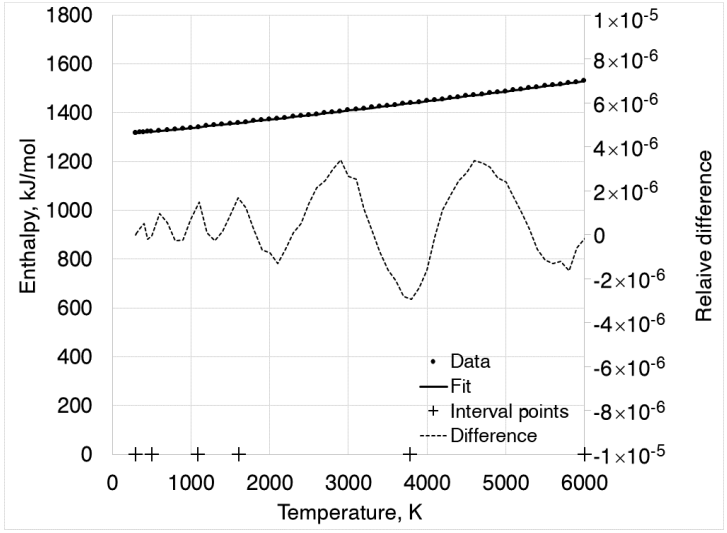


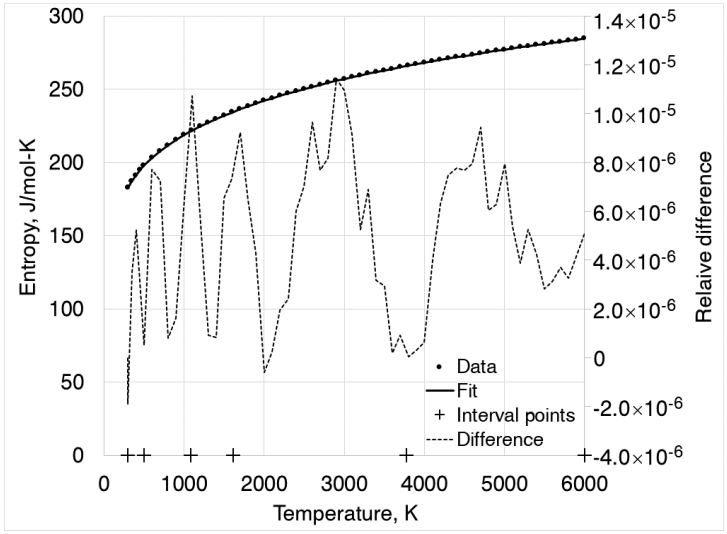


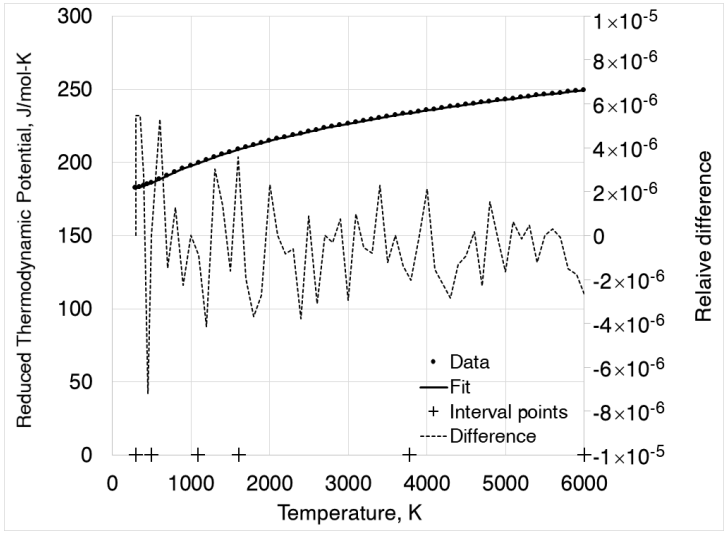
Hydroxyl, Ion (OH^+) – H1O1+(g)

MW = 17.0068 g/mol – $\Delta_f H^o$(298.15 K) = 1317.123 kJ/mol – S^o(298.15 K) = 182.776 J/mol-K – p^o = 0.1 MPa



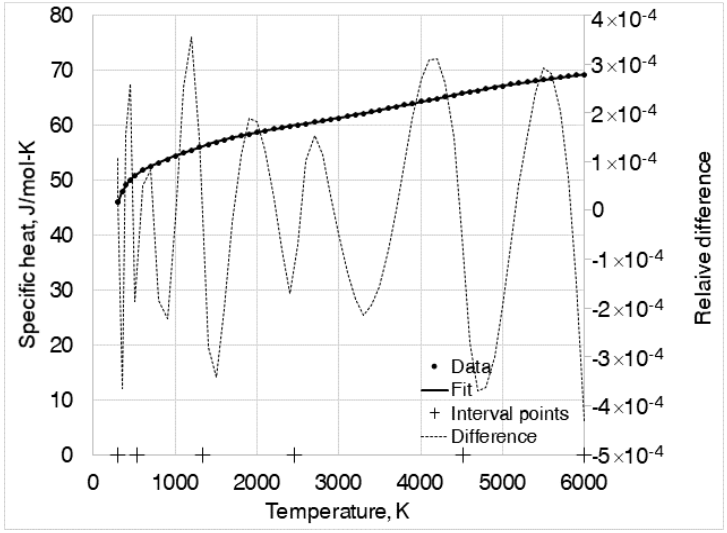


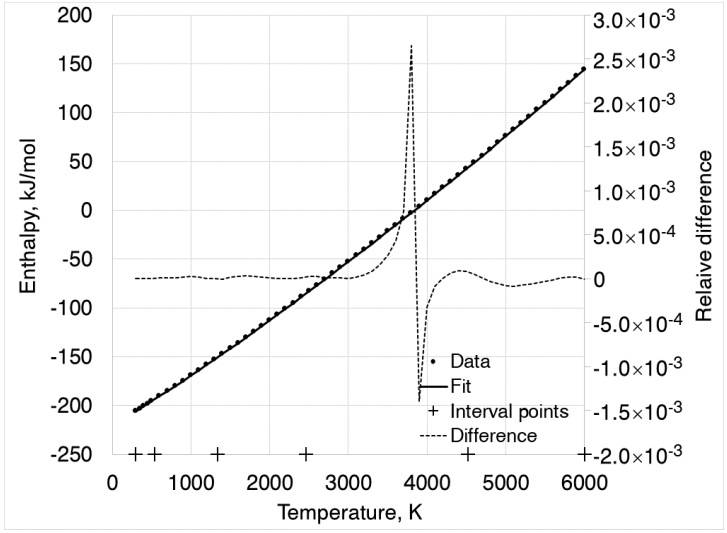


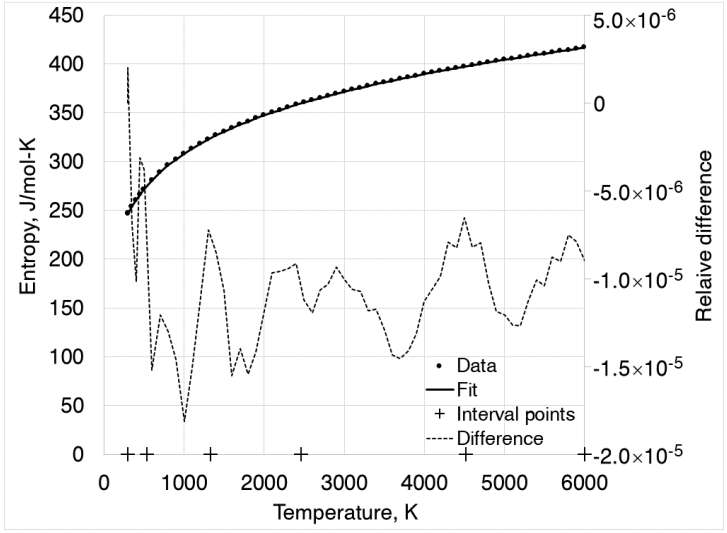


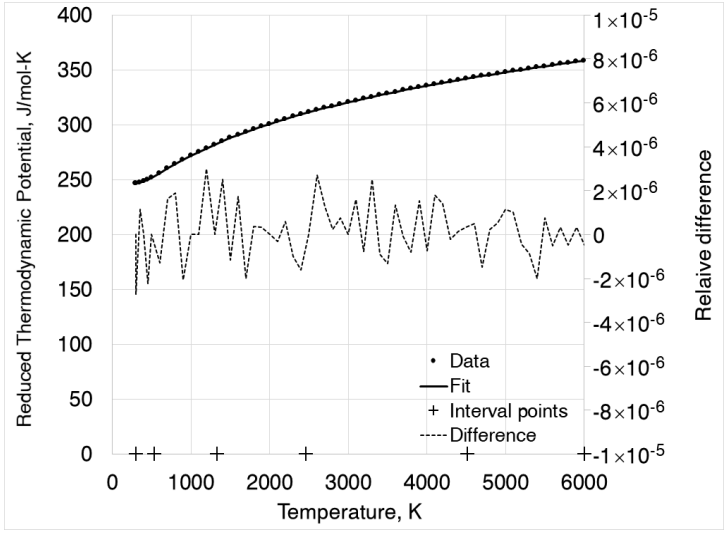
Strontium Hydroxide (SrOH) – H1O1Sr1(g)

MW = 104.63 g/mol – $\Delta_f H^o$(298.15 K) = -205.518 kJ/mol – S^o(298.15 K) = 246.511 J/mol-K – p^o = 0.1 MPa



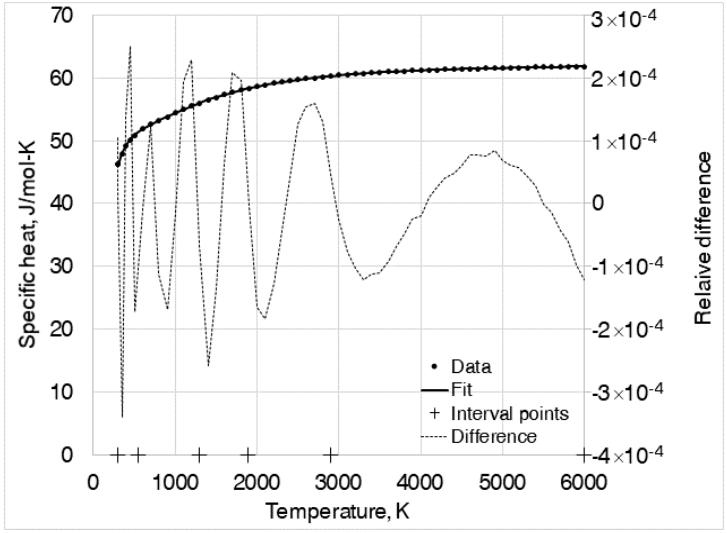


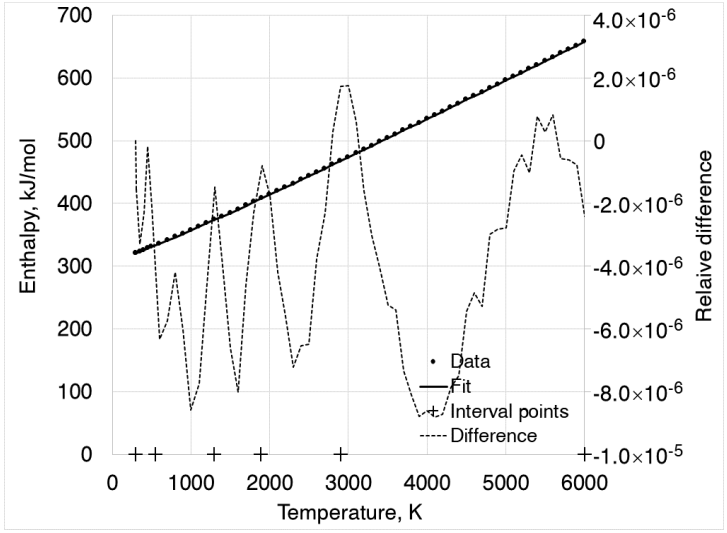


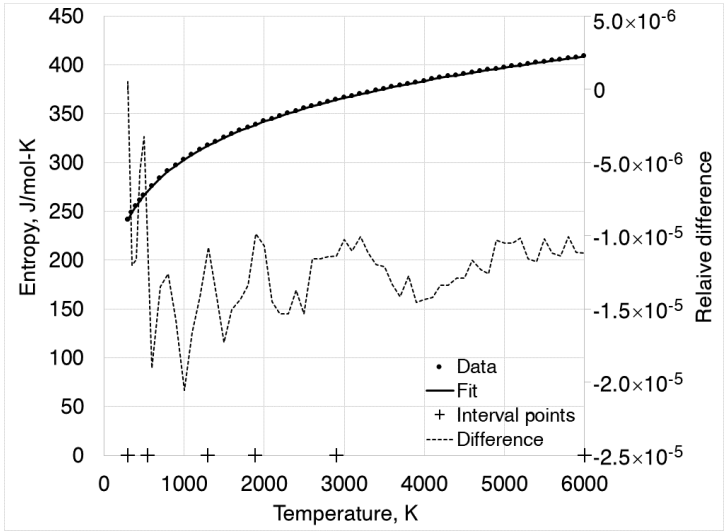


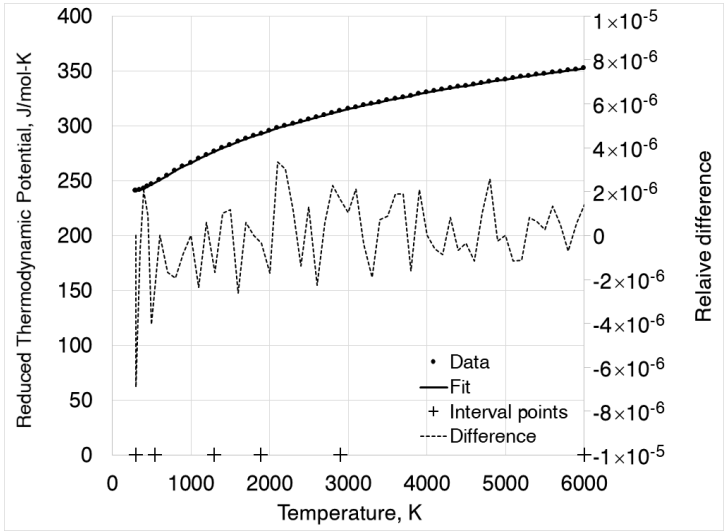
Strontium Hydroxide, Ion ($SrOH^+$) – H1O1Sr1+(g)

MW = 104.63 g/mol – $\Delta_f H^o$(298.15 K) = 320.8 kJ/mol – S^o(298.15 K) = 241.049 J/mol-K – p^o = 0.1 MPa



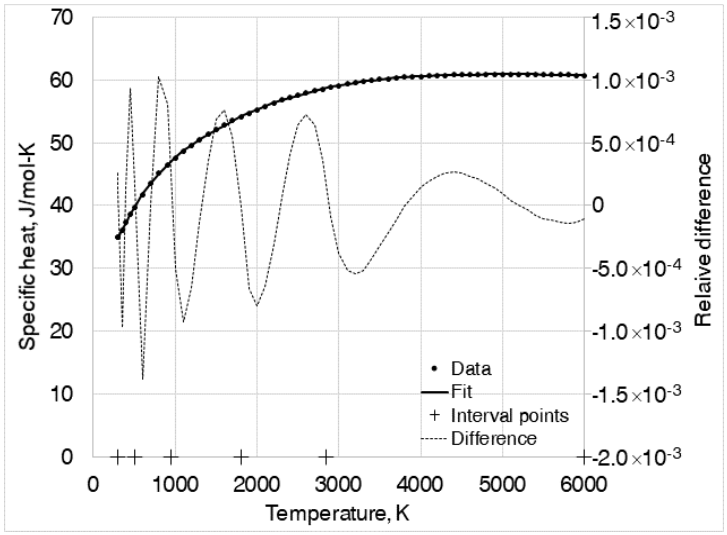


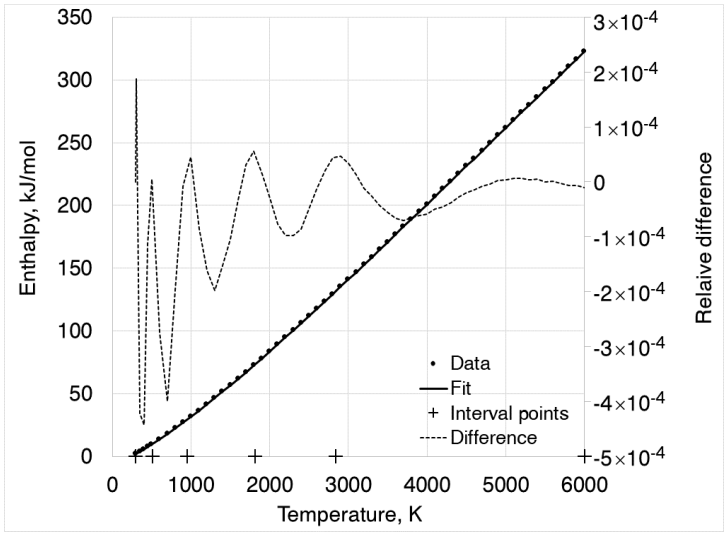


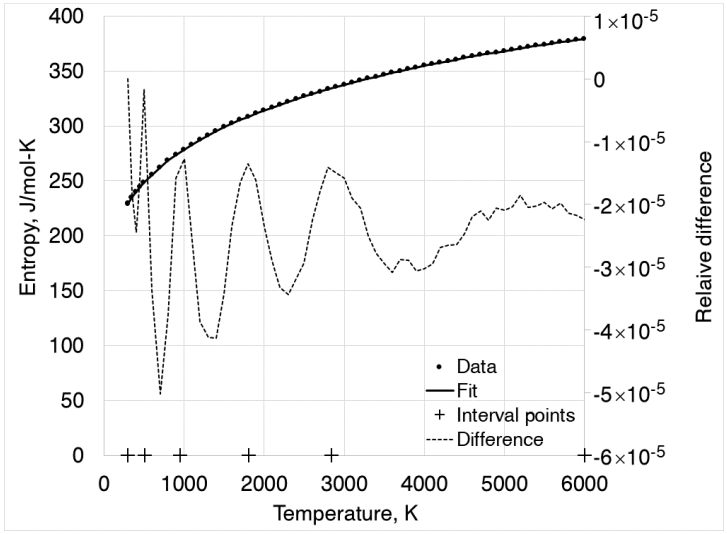


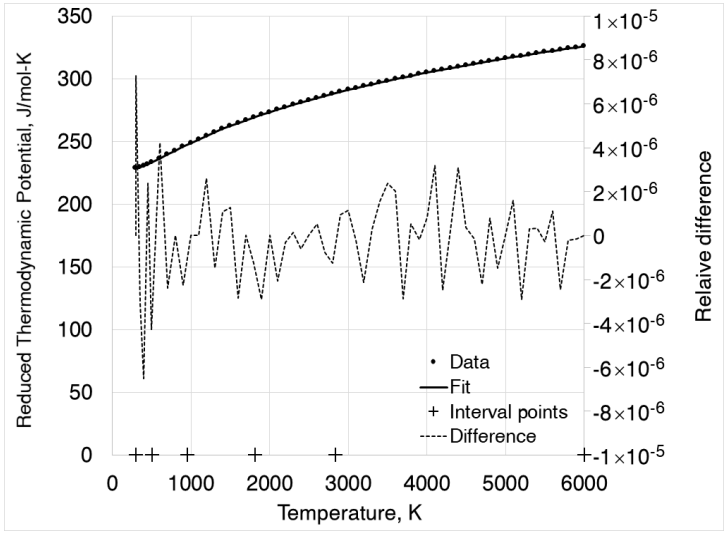
Hydroperoxo (HOO) – H1O2(g)

MW = 33.0067 g/mol – $\Delta_f H^o$(298.15 K) = 2.092 kJ/mol – S^o(298.15 K) = 229.106 J/mol-K – p^o = 0.1 MPa



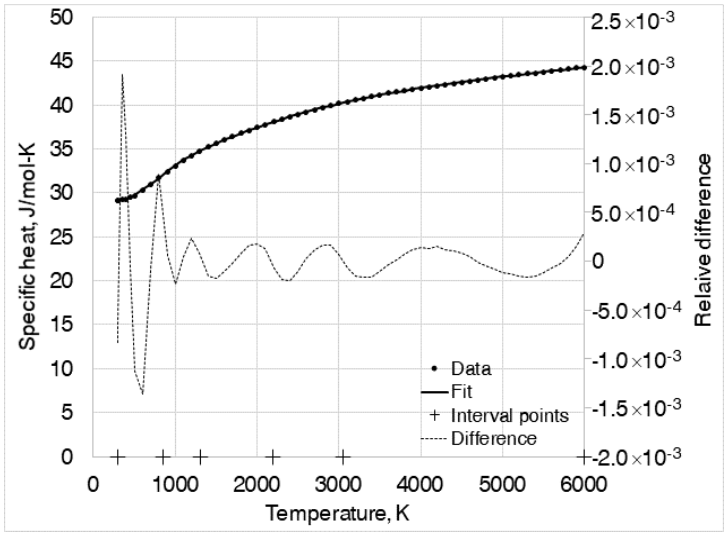


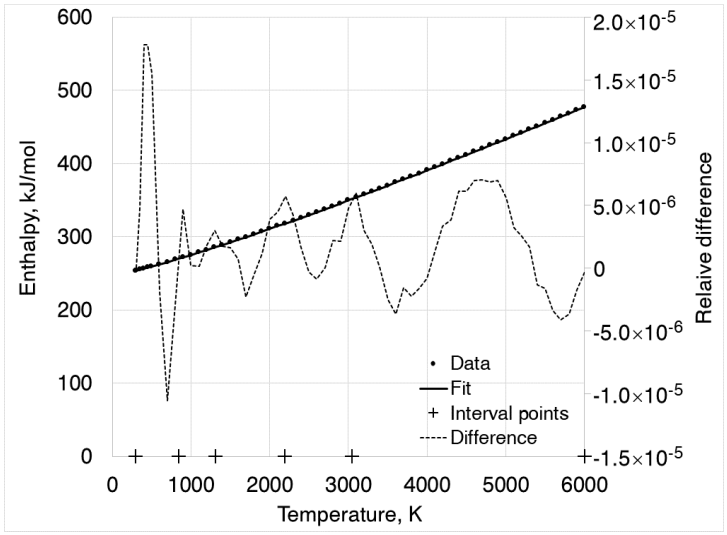


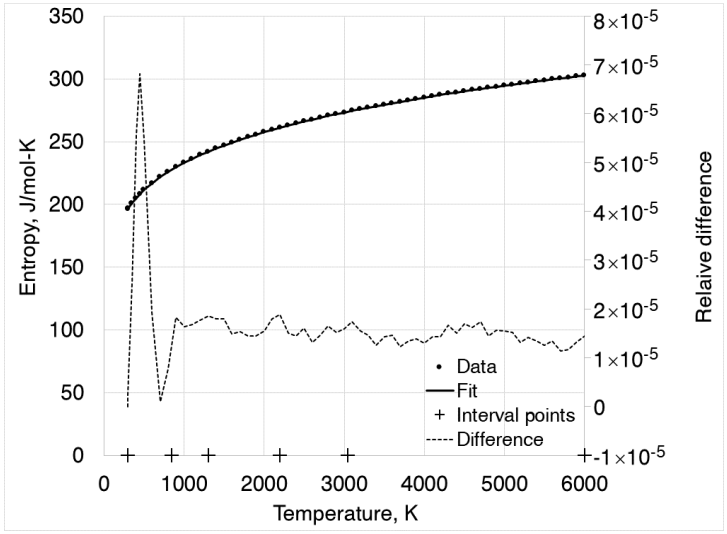


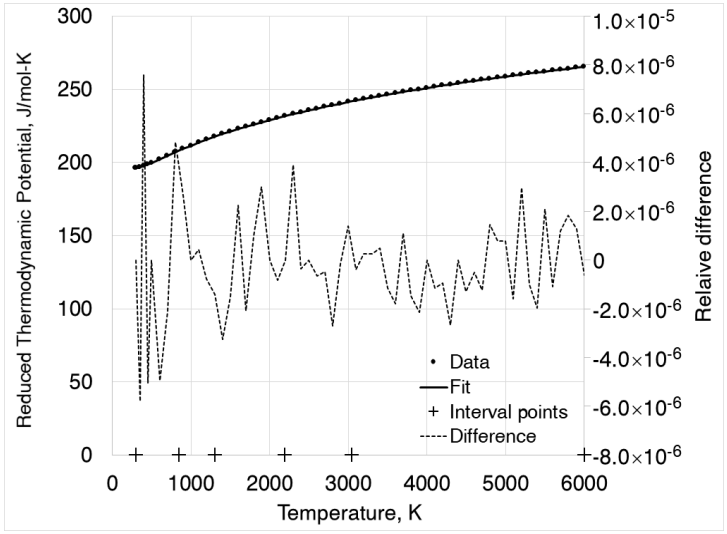
Phosphinidene (PH) – H1P1(g)

MW = 31.9817 g/mol – $\Delta_f H^o$(298.15 K) = 253.55 kJ/mol – S^o(298.15 K) = 196.344 J/mol-K – p^o = 0.1 MPa



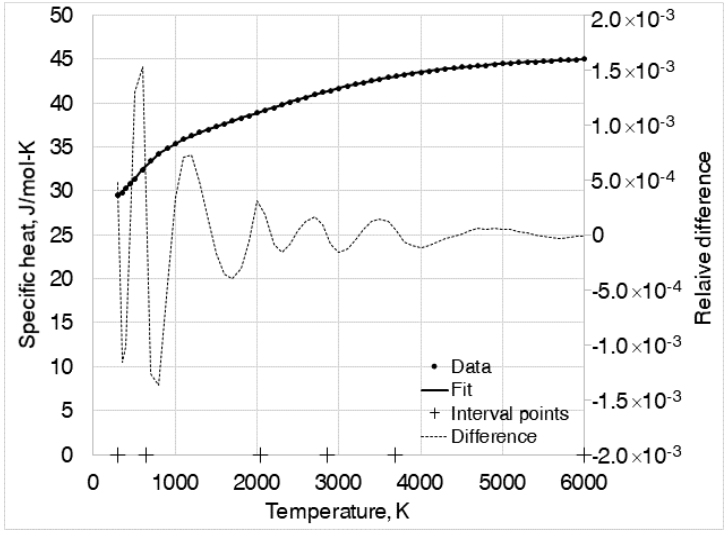


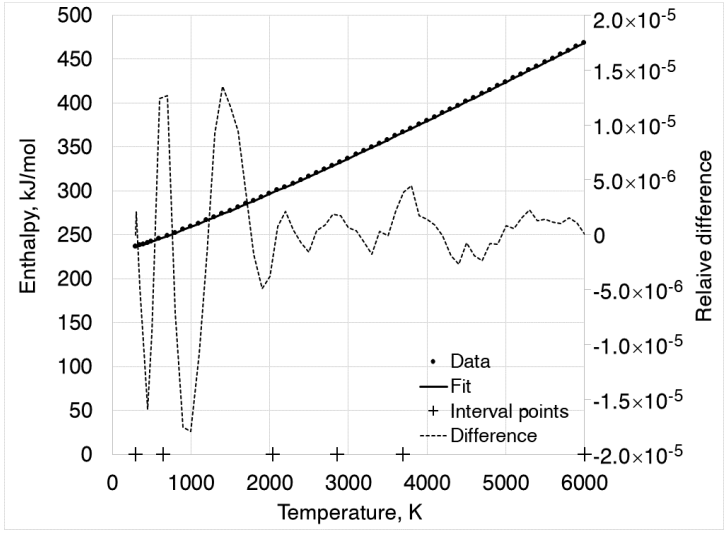


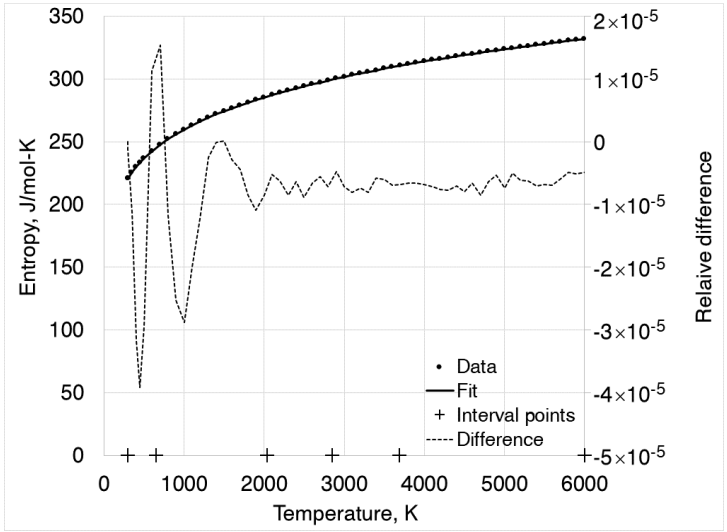


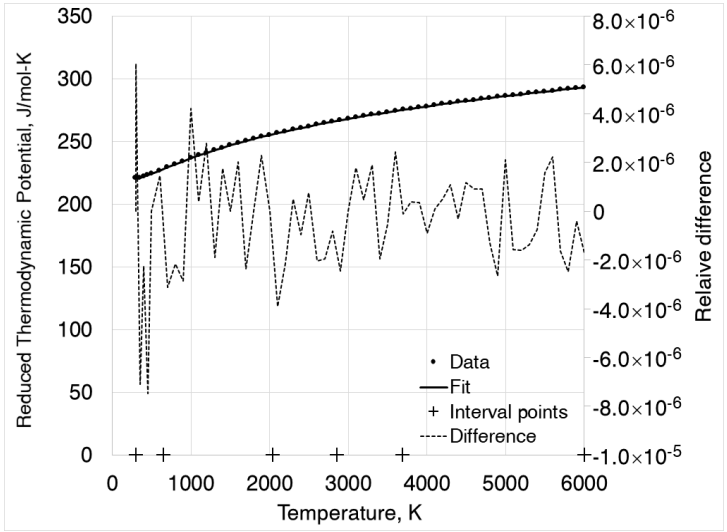
Lead Hydride (PbH) – H1Pb1(g)

MW = 208.2 g/mol – $\Delta_f H^o$(298.15 K) = 236.187 kJ/mol – S^o(298.15 K) = 220.756 J/mol-K – p^o = 0.1 MPa



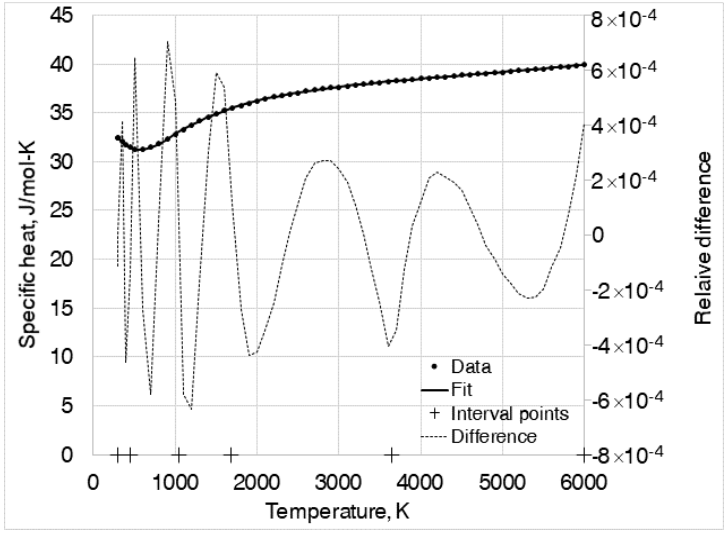


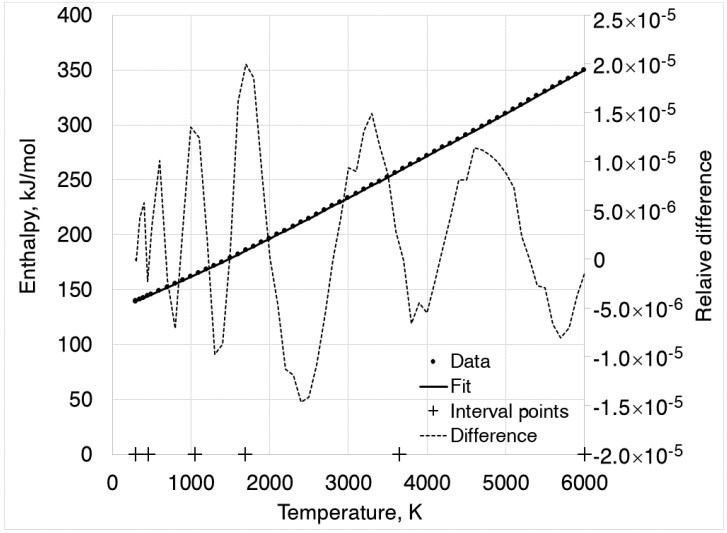


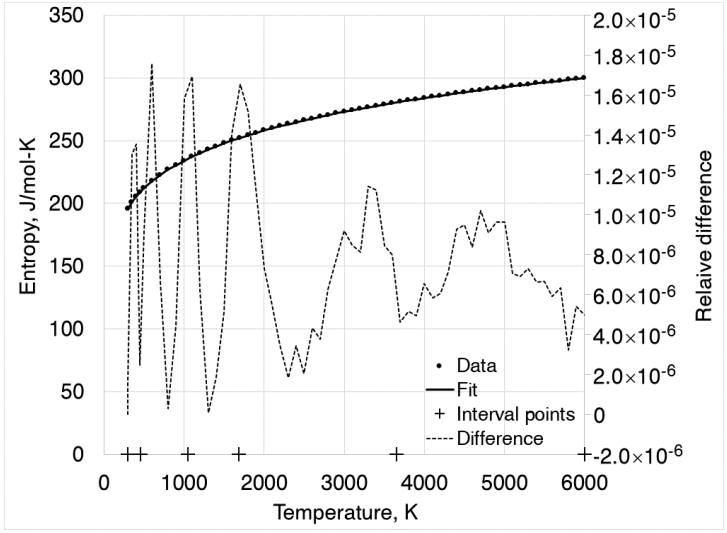


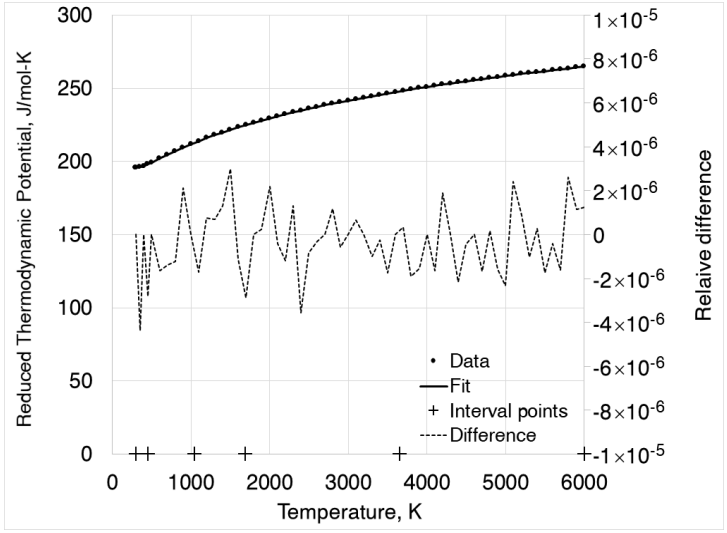
Mercapto (HS) – H1S1(g)

MW = 33.07 g/mol – $\Delta_f H^o$(298.15 K) = 139.327 kJ/mol – S^o(298.15 K) = 195.628 J/mol-K – p^o = 0.1 MPa



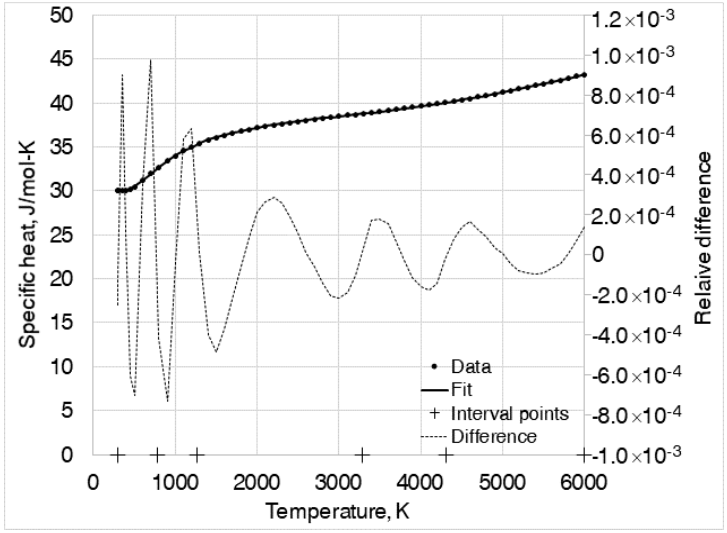


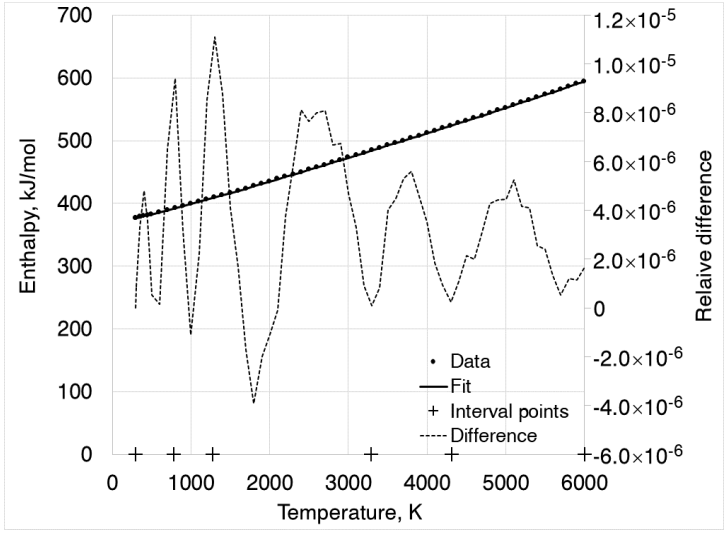


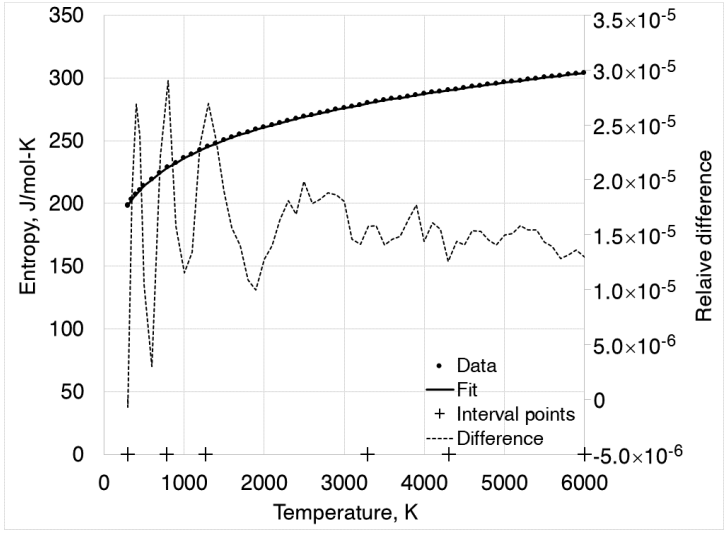


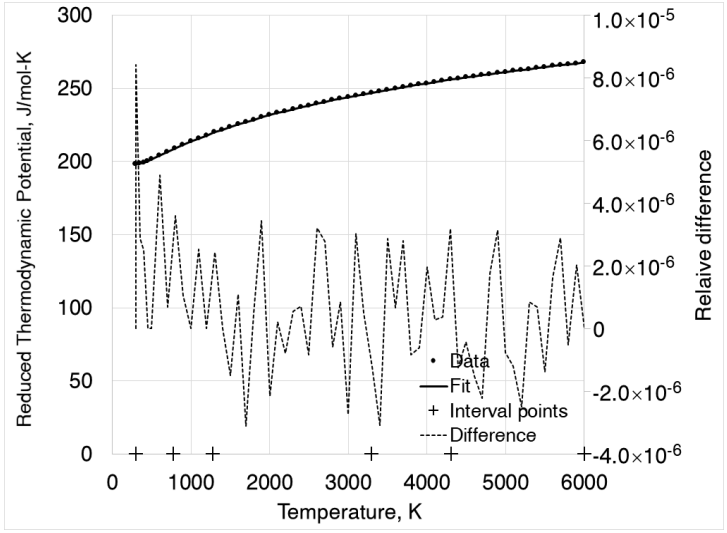
Silylidyne (SiH) – H1Si1(g)

MW = 29.0934 g/mol – $\Delta_f H^o$(298.15 K) = 376.662 kJ/mol – S^o(298.15 K) = 198.04 J/mol-K – p^o = 0.1 MPa



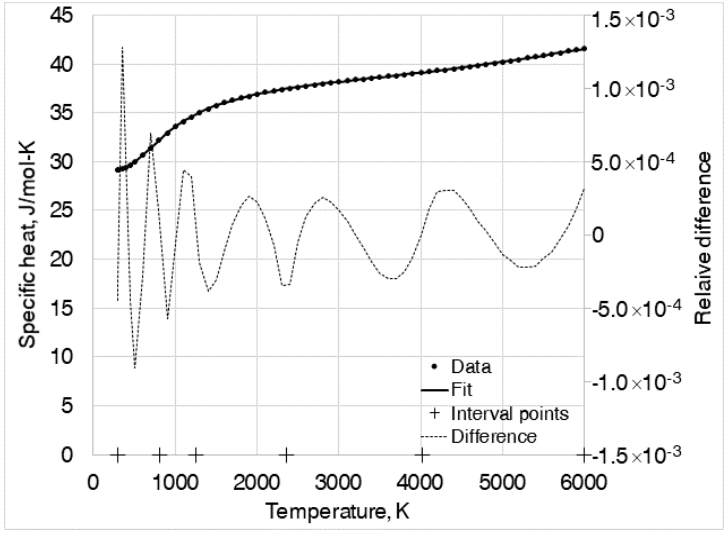


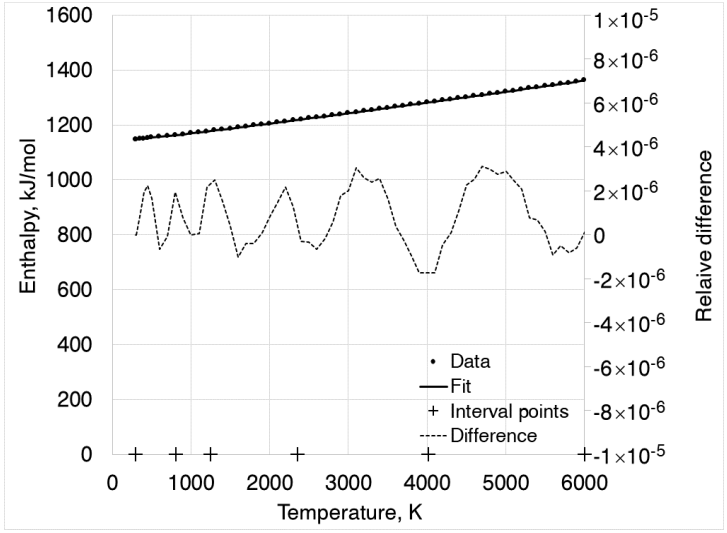


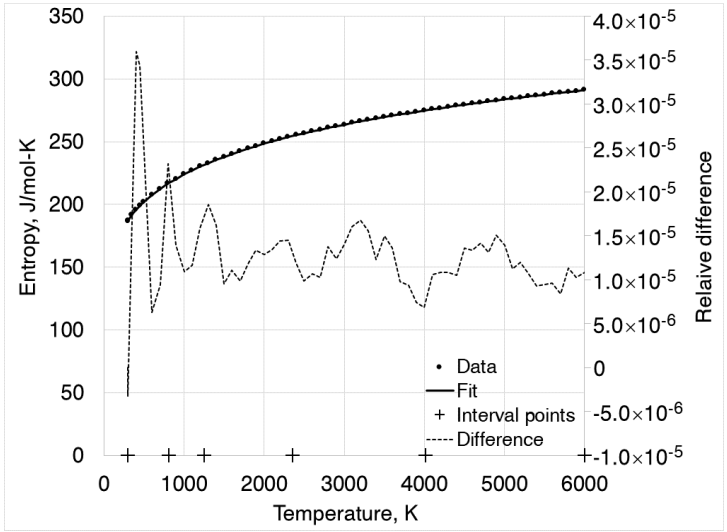


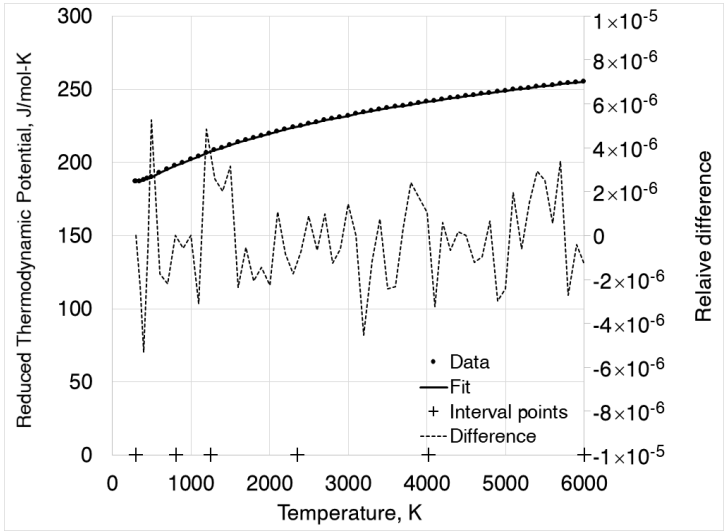
Silylidyne, Ion (SiH^+) – H1Si1+(g)

MW = 29.0929 g/mol – $\Delta_f H^o$(298.15 K) = 1147.671 kJ/mol – S^o(298.15 K) = 186.789 J/mol-K – p^o = 0.1 MPa



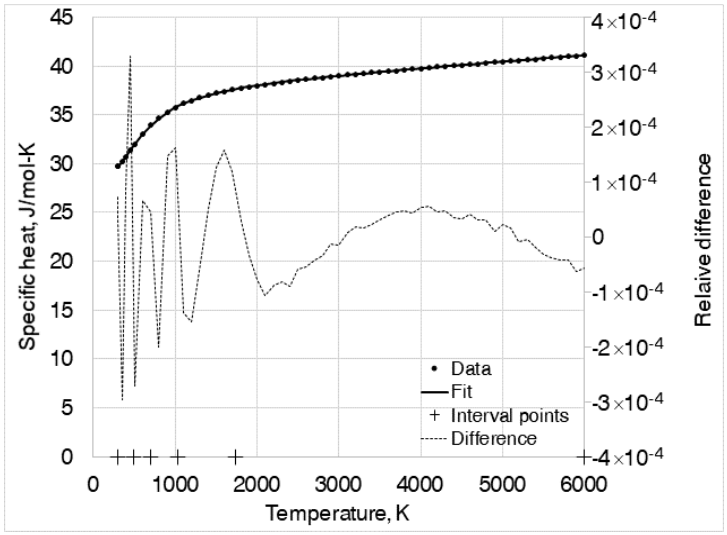


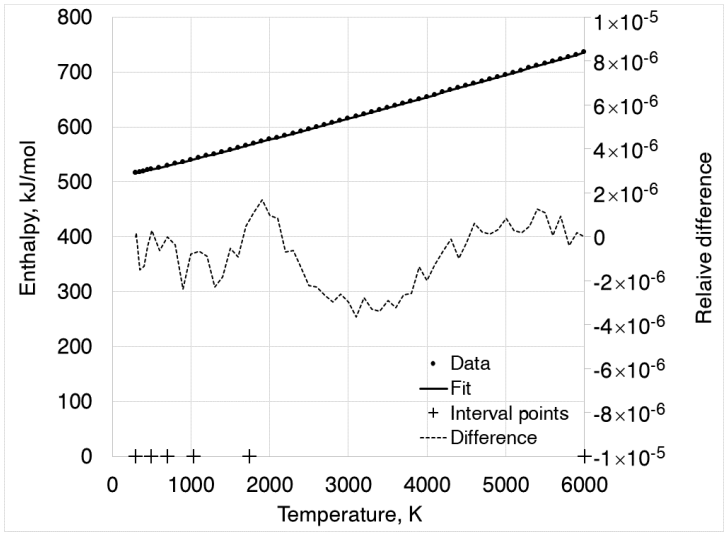


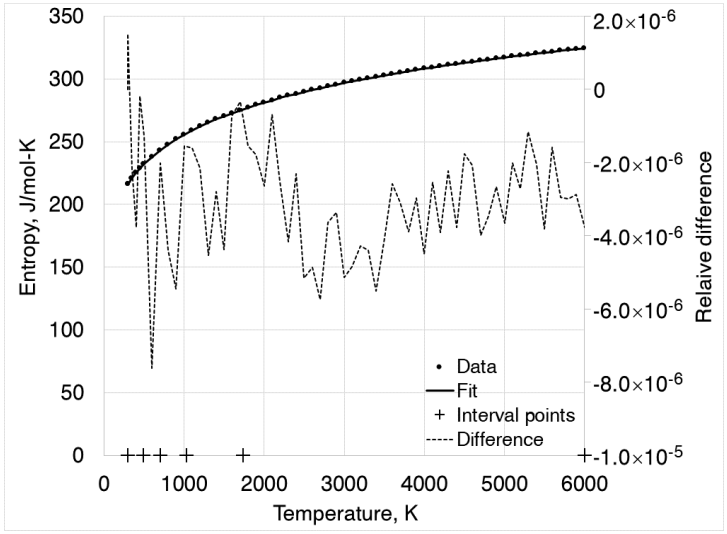


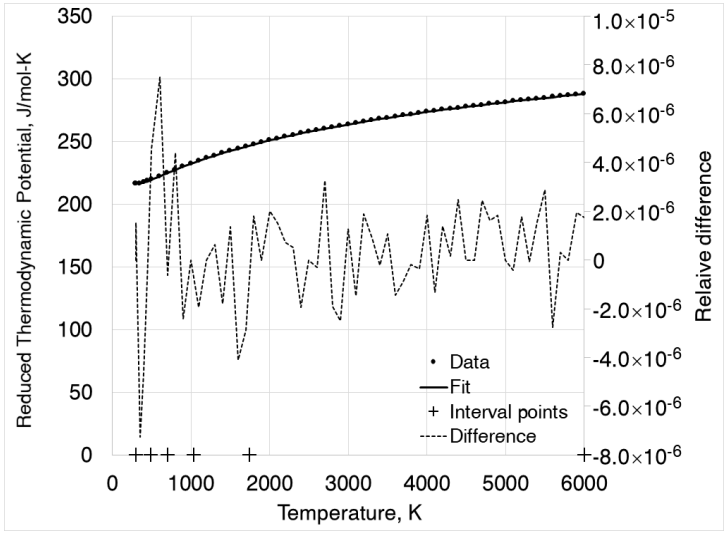
Zirconium Hydride (ZrH) – H1Zr1(g)

MW = 92.23 g/mol – $\Delta_f H^o$(298.15 K) = 516.306 kJ/mol – S^o(298.15 K) = 216.158 J/mol-K – p^o = 0.1 MPa



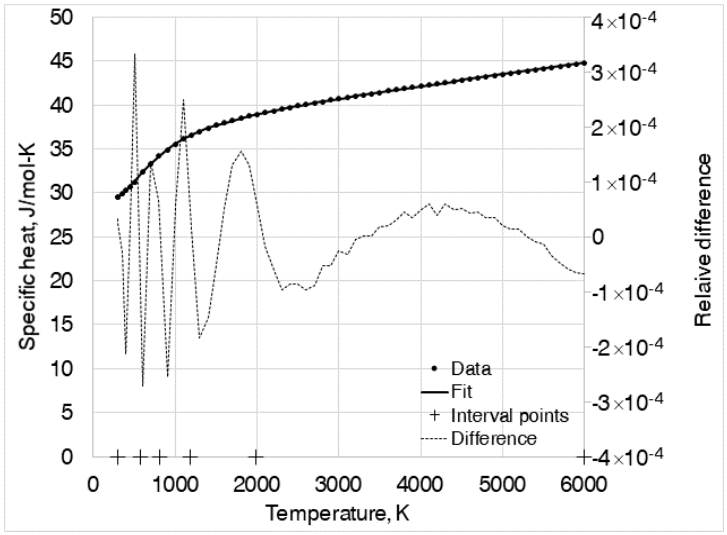


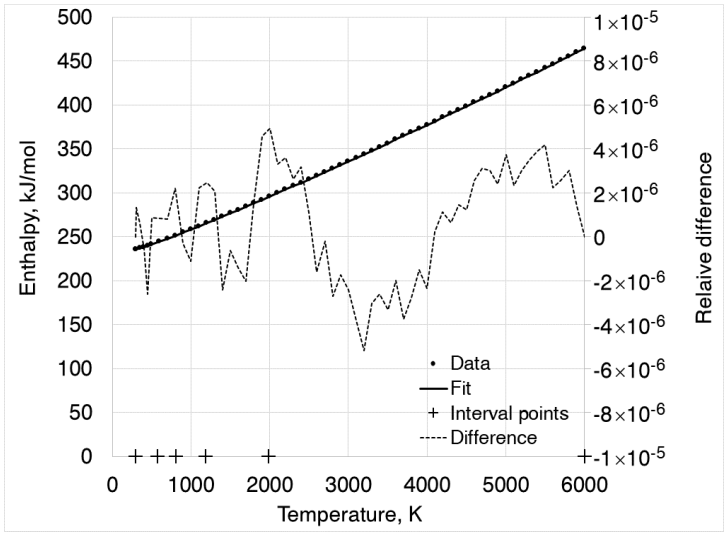


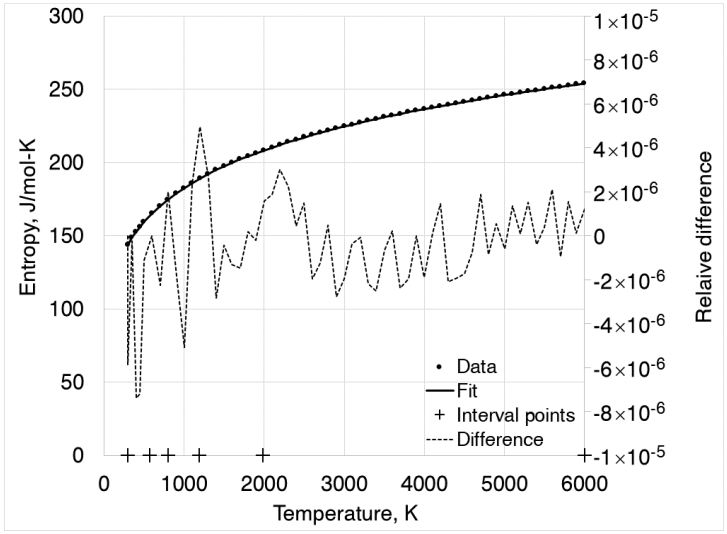


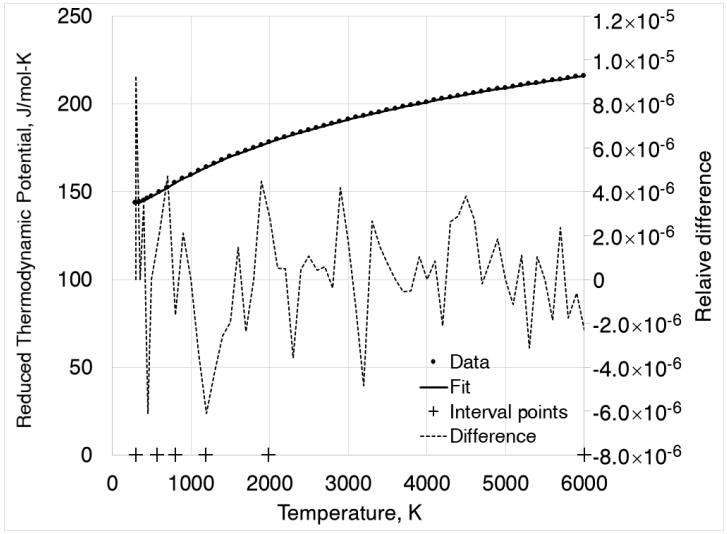
Hydrogen, Ion (H_2^-) – H2-(g)

MW = 2.01643 g/mol – $\Delta_f H^o$(298.15 K) = 235.373 kJ/mol – S^o(298.15 K) = 143.746 J/mol-K – p^o = 0.1 MPa



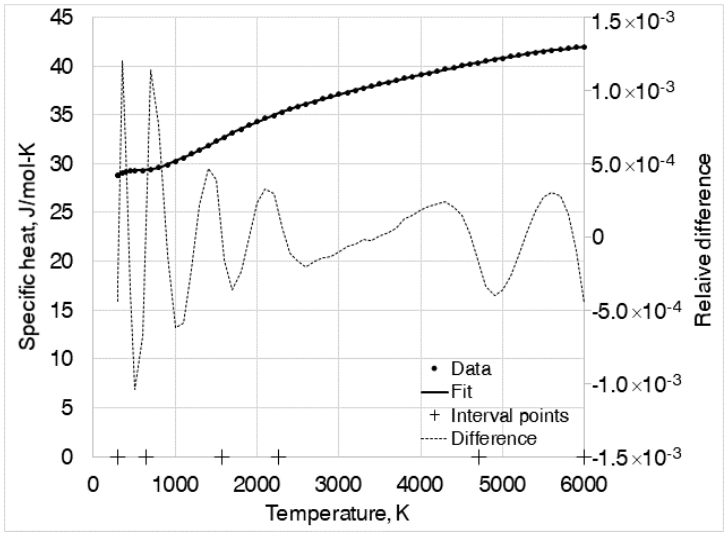


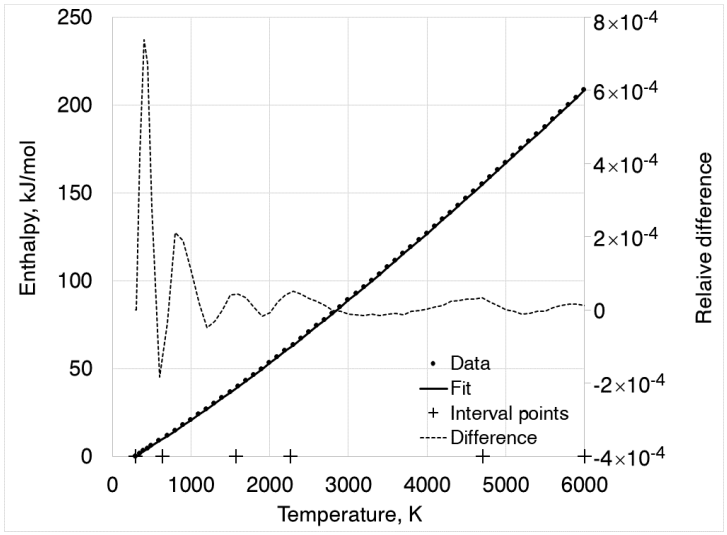


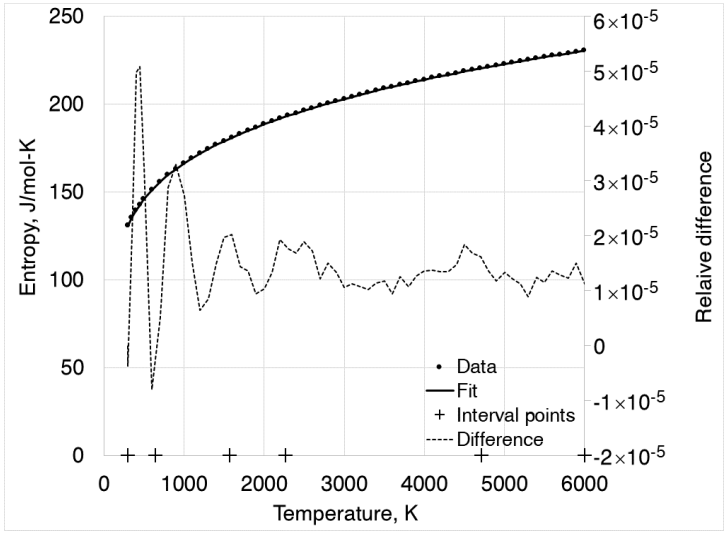


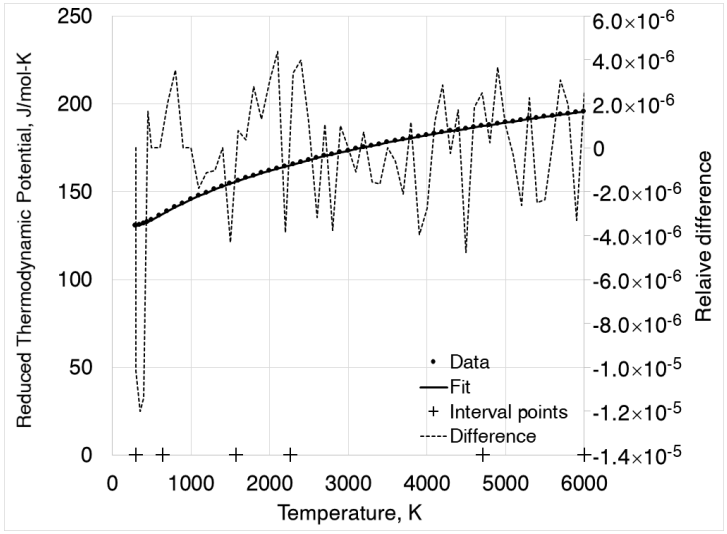
Hydrogen (H_2) – H2(ref)

MW = 2.01588 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 130.68 J/mol-K – p^o = 0.1 MPa



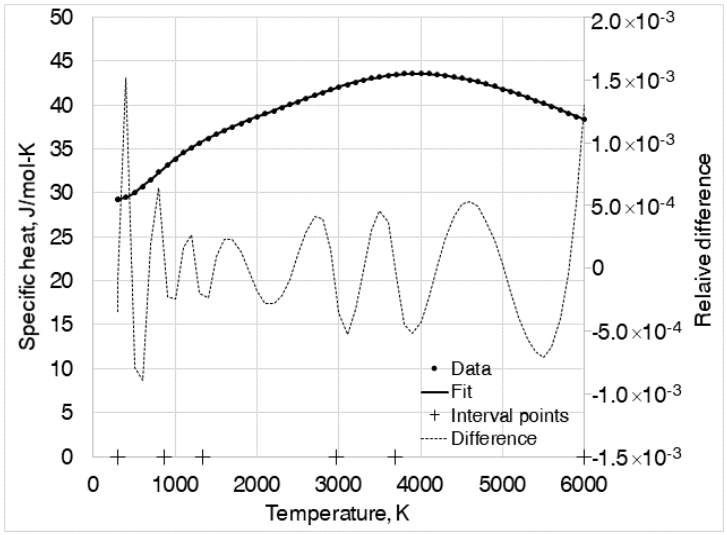


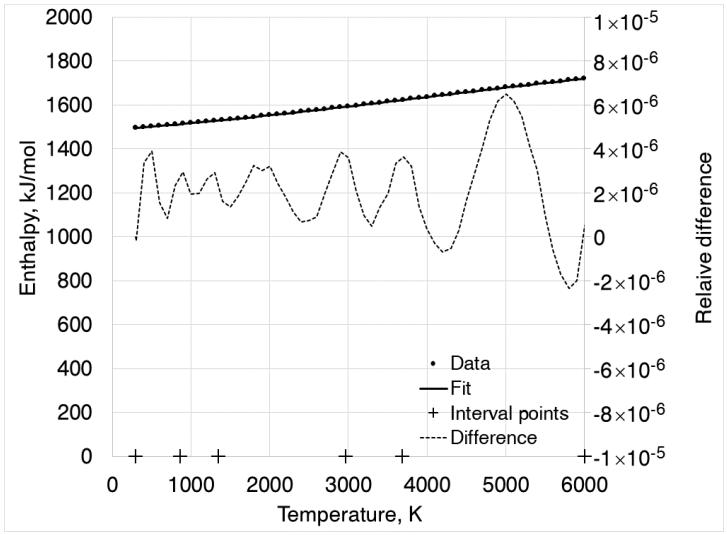


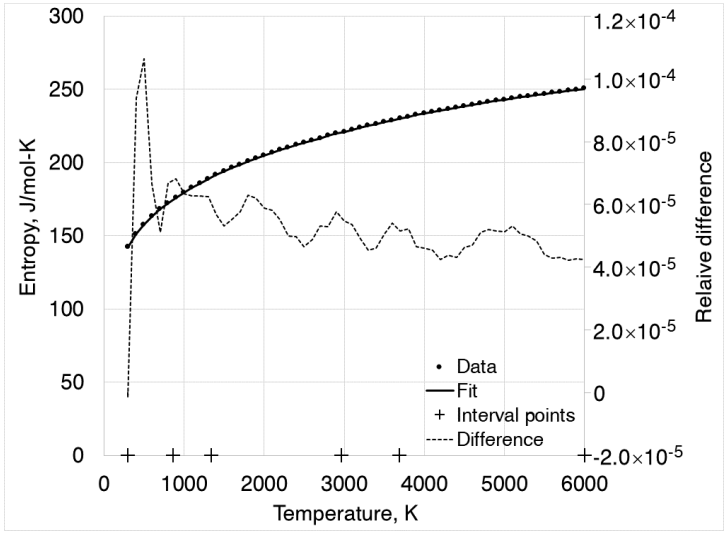


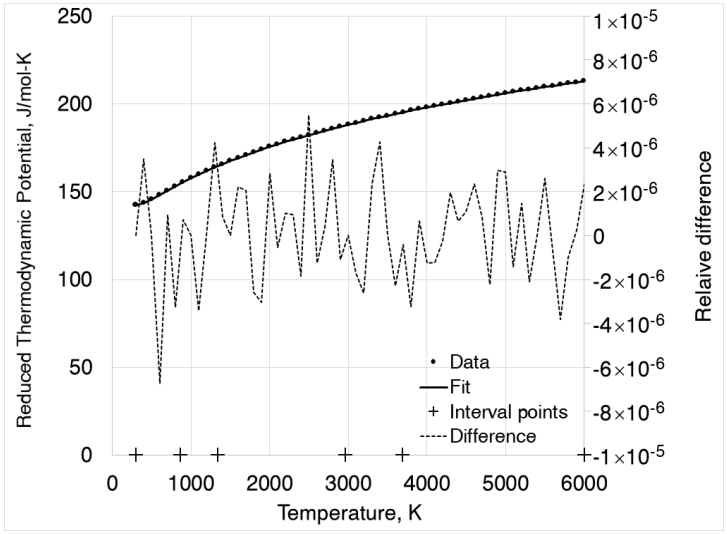
Hydrogen, Ion (H_2^+) – H2+(g)

MW = 2.01533 g/mol – $\Delta_f H^o$(298.15 K) = 1494.671 kJ/mol – S^o(298.15 K) = 142.37 J/mol-K – p^o = 0.1 MPa



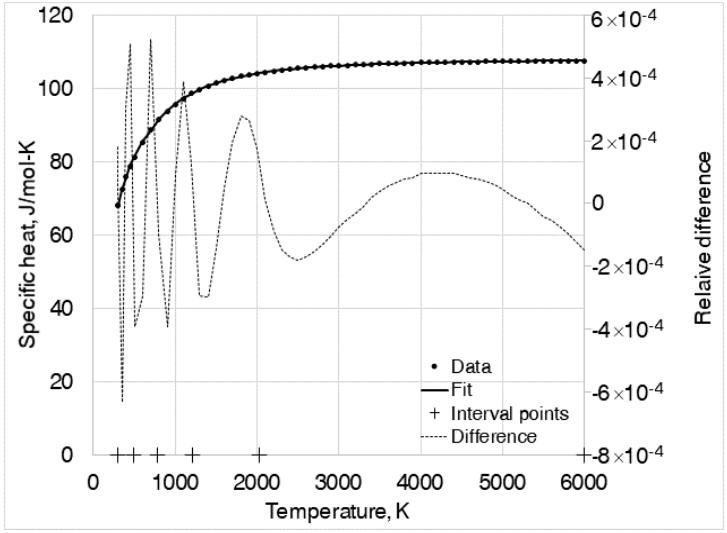


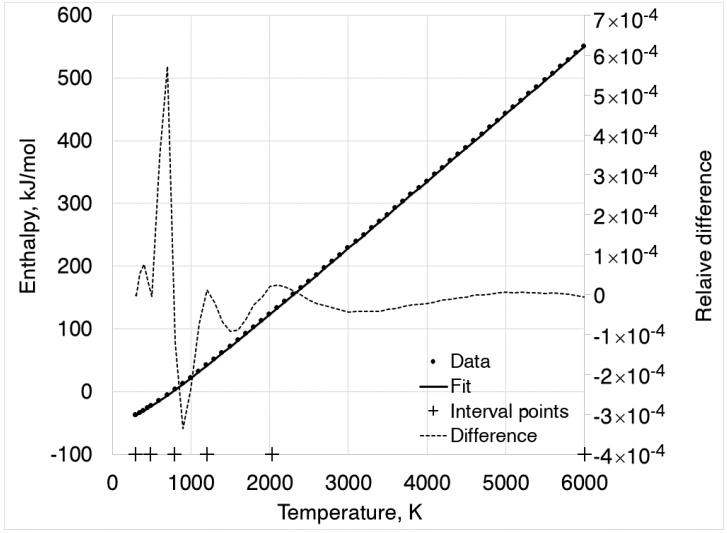


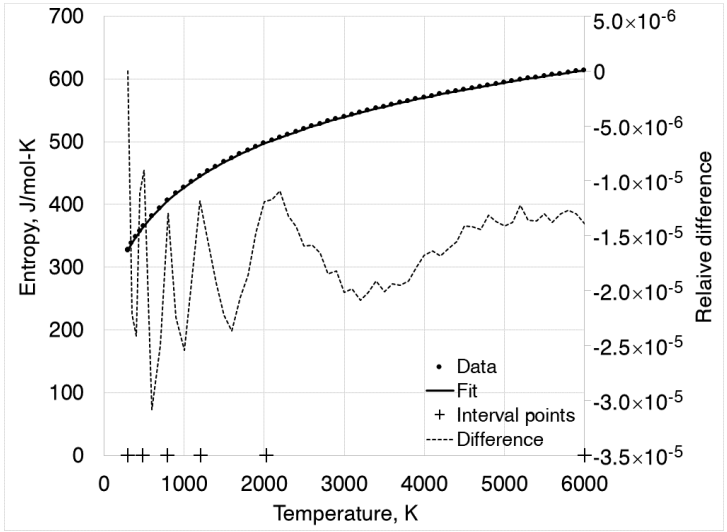


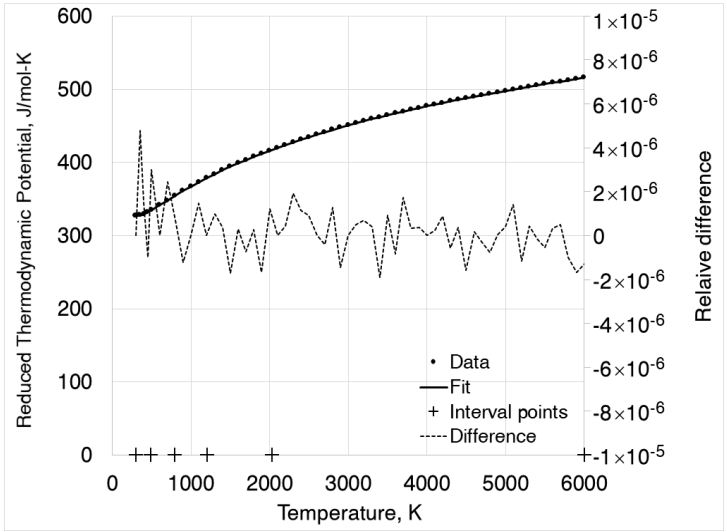
Diiodosilane (SiH_2I_2) – H2I2Si1(g)

MW = 283.9104 g/mol – $\Delta_f H^o$(298.15 K) = -38.074 kJ/mol – S^o(298.15 K) = 326.797 J/mol-K – p^o = 0.1 MPa



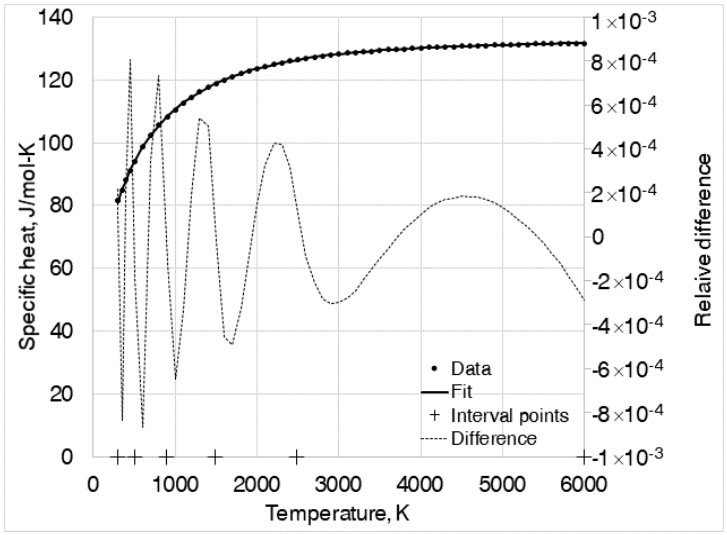


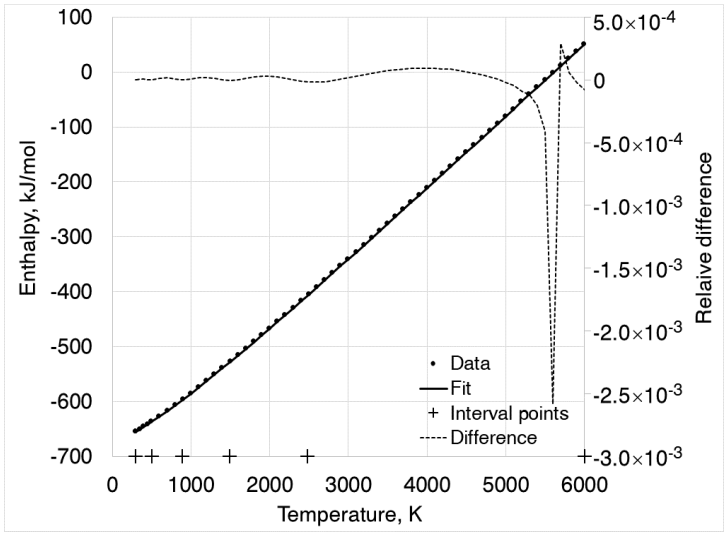


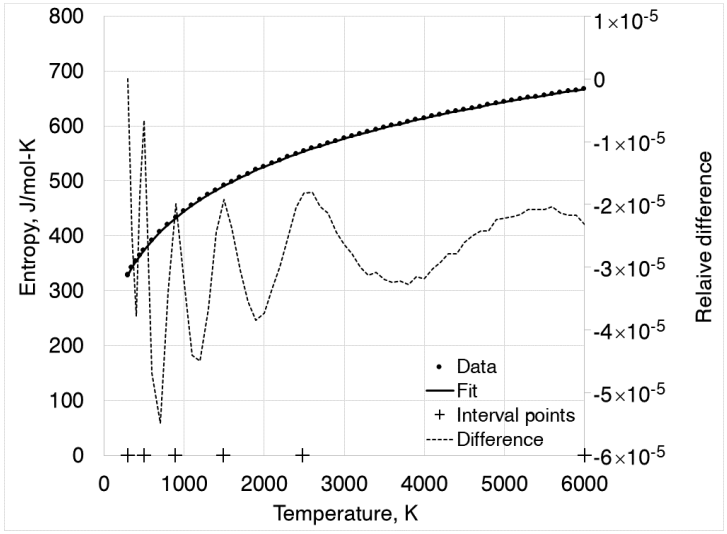


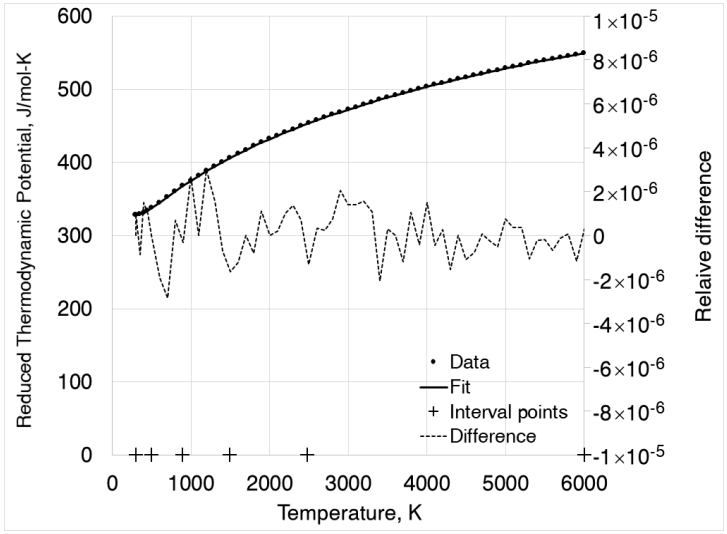
Potassium Hydroxide ($(KOH)_2$) – H2K2O2(g)

MW = 112.2113 g/mol – $\Delta_f H^o$(298.15 K) = -654.796 kJ/mol – S^o(298.15 K) = 327.996 J/mol-K – p^o = 0.1 MPa



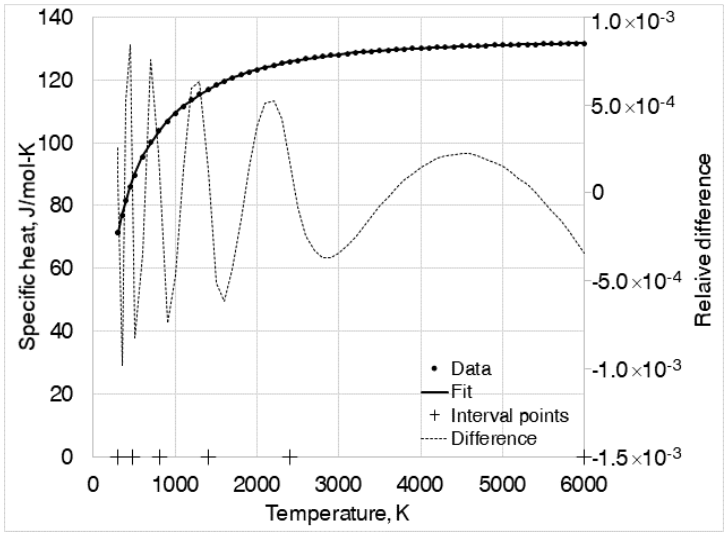


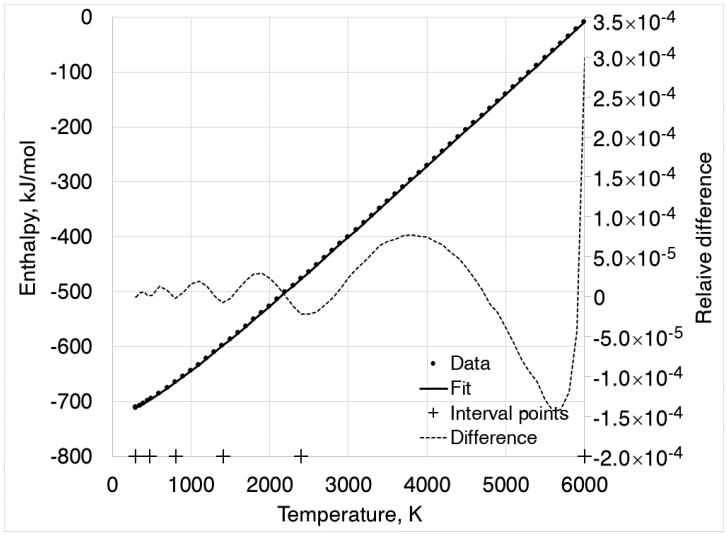


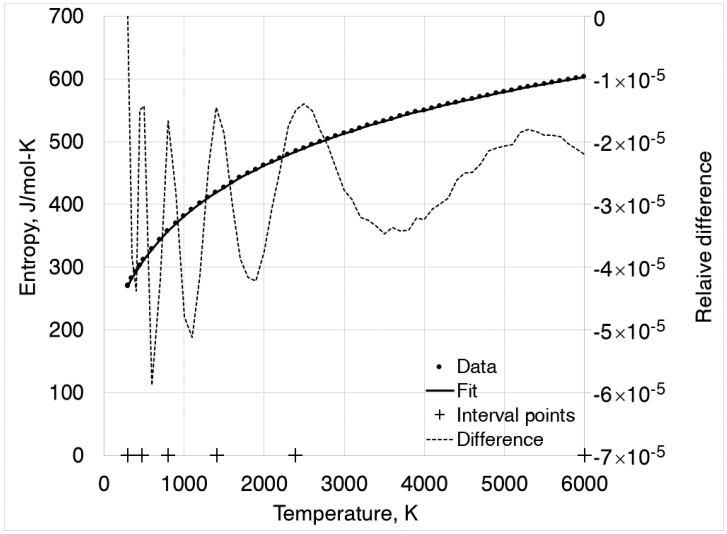


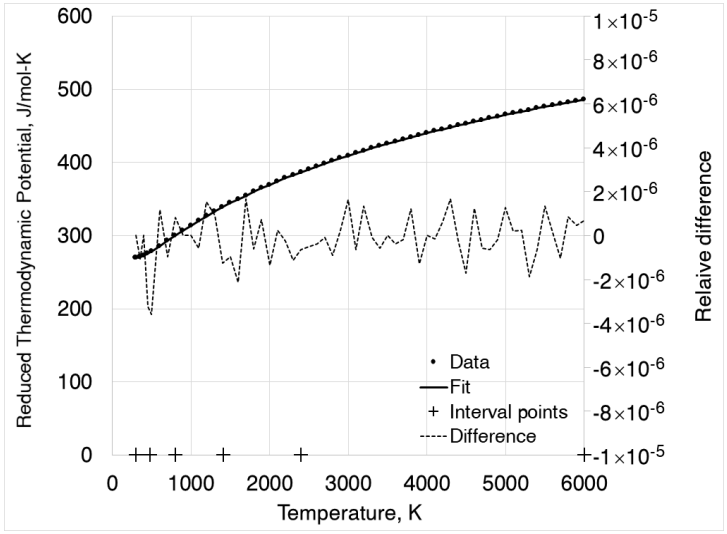
Lithium Hydroxide ($(LiOH)_2$) – H2Li2O2(g)

MW = 47.897 g/mol – $\Delta_f H^o$(298.15 K) = -711.28 kJ/mol – S^o(298.15 K) = 269.794 J/mol-K – p^o = 0.1 MPa



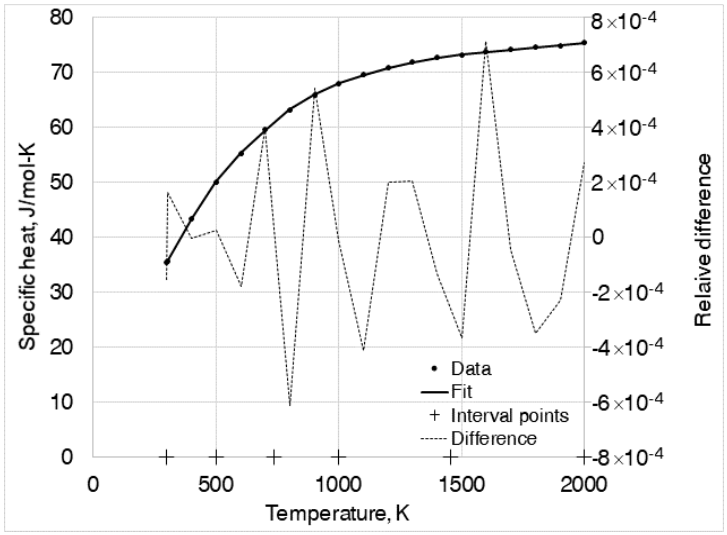


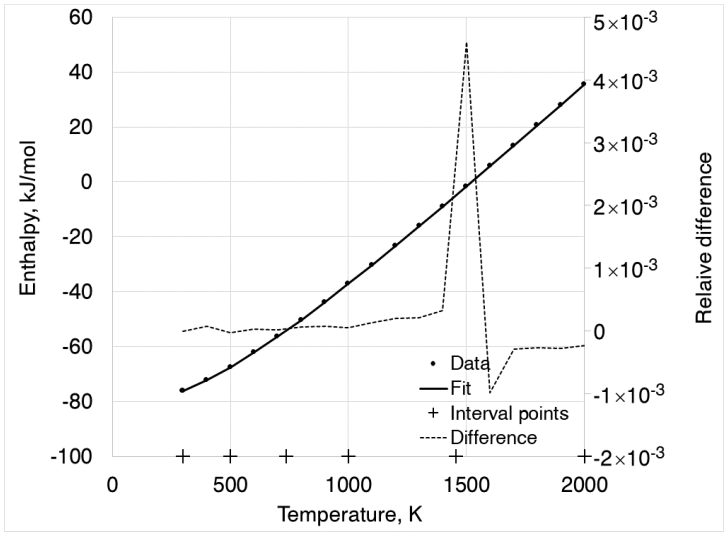


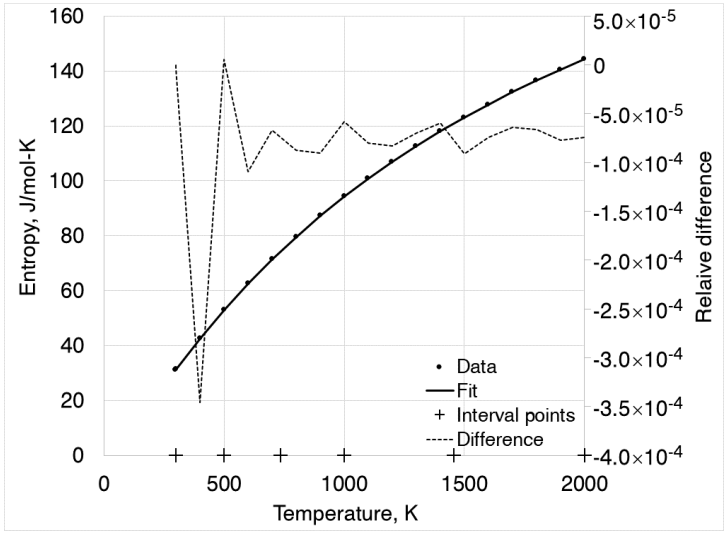


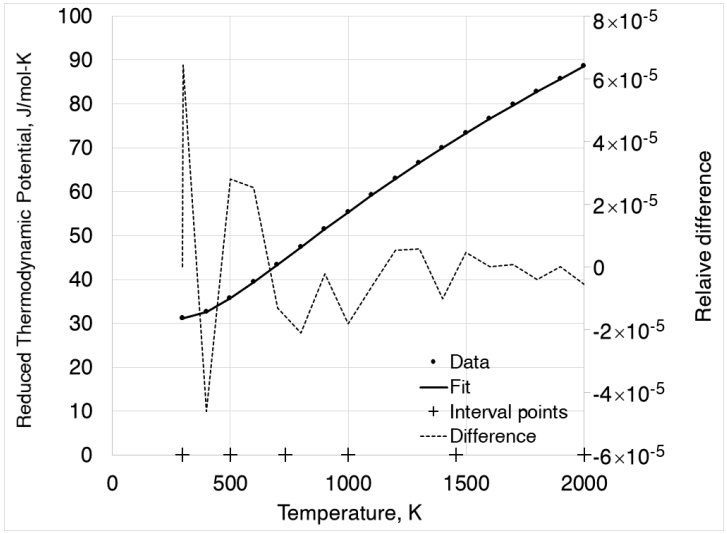
Magnesium Hydride (MgH_2) – H2Mg1(cr)

MW = 26.321 g/mol – $\Delta_f H^o$(298.15 K) = -76.149 kJ/mol – S^o(298.15 K) = 31.091 J/mol-K – p^o = 0.1 MPa



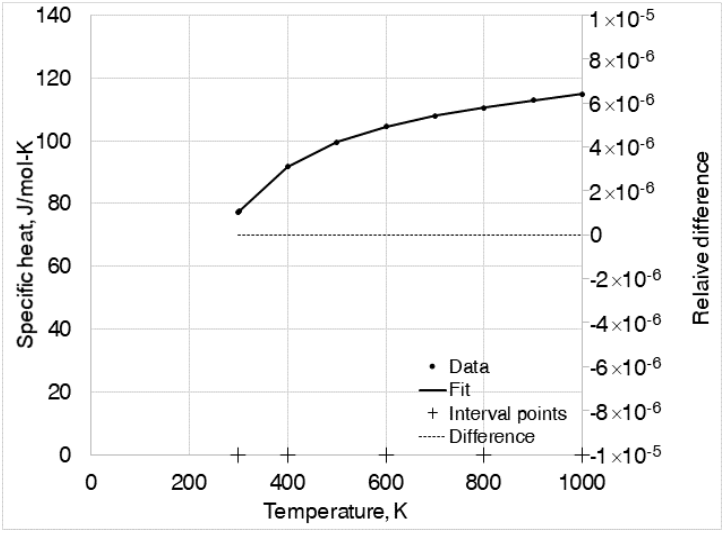


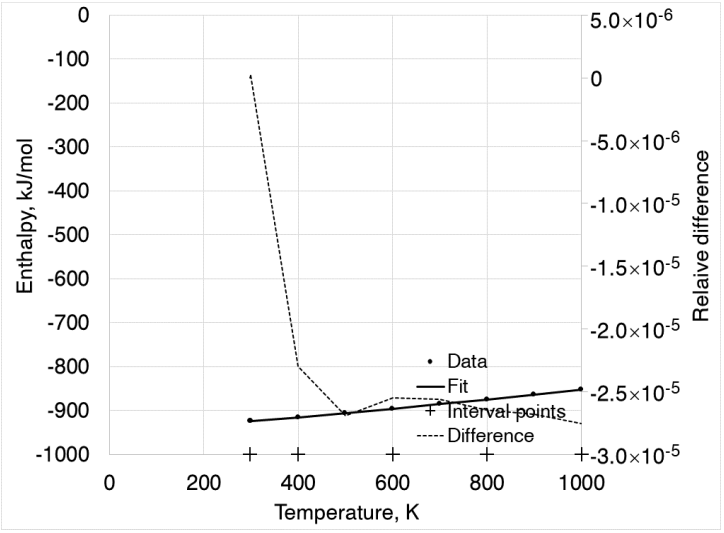


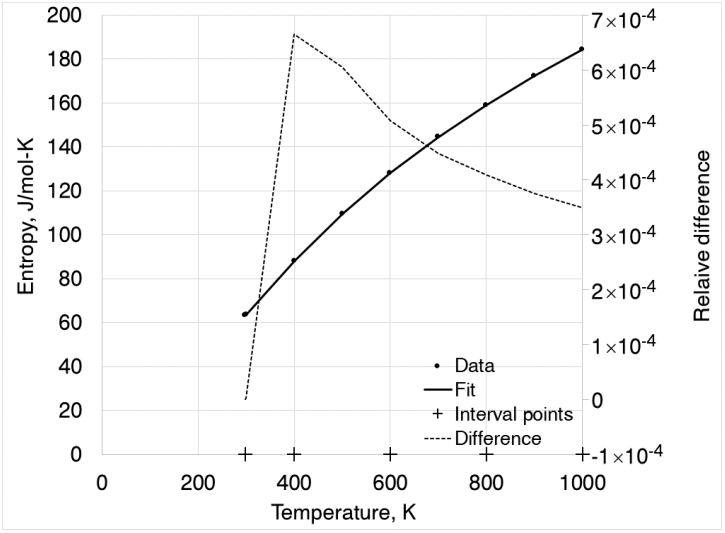


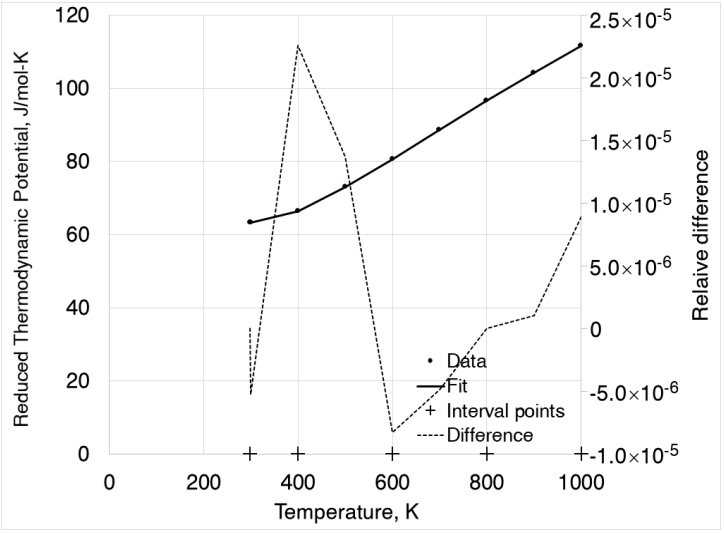
Magnesium Hydroxide ($Mg(OH)_2$) – H2Mg1O2(cr)

MW = 58.32 g/mol – $\Delta_f H^o$(298.15 K) = -924.664 kJ/mol – S^o(298.15 K) = 63.242 J/mol-K – p^o = 0.1 MPa



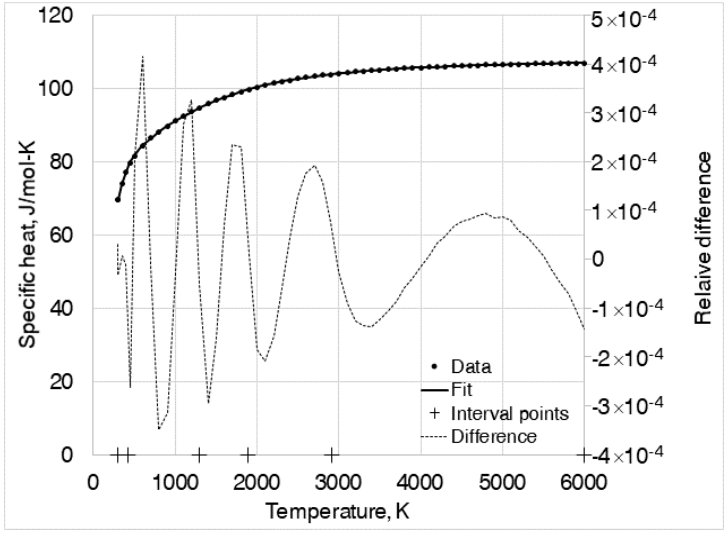


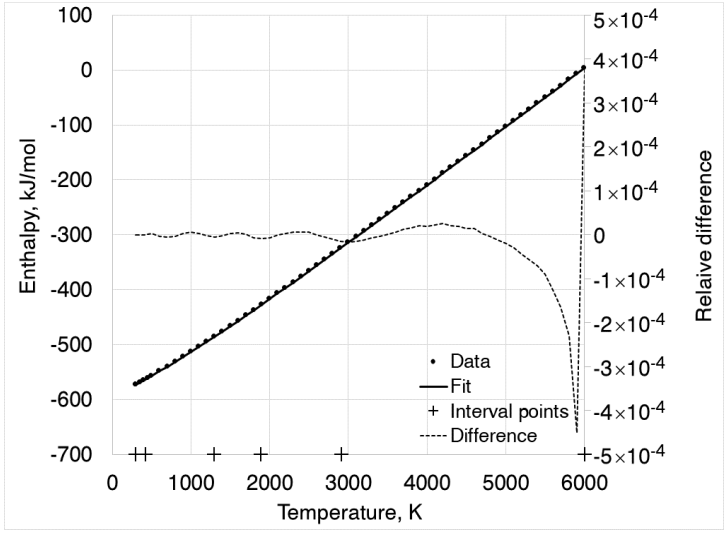


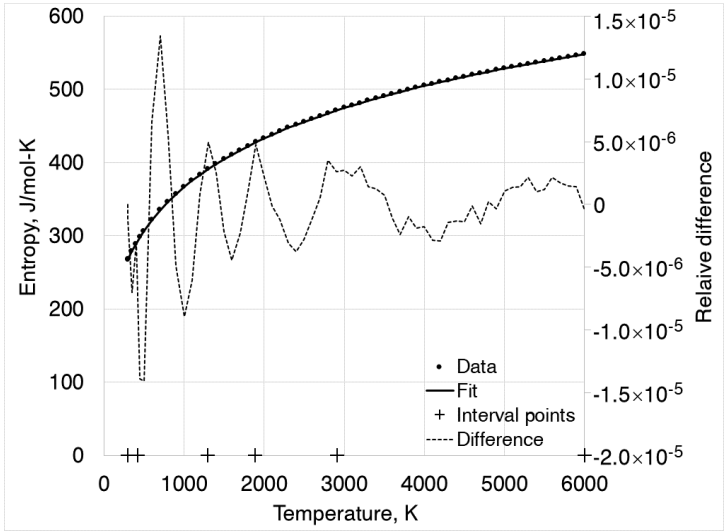


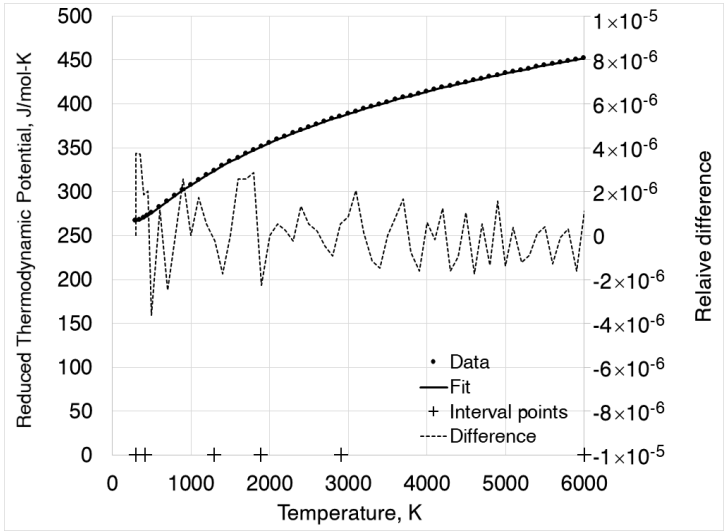
Magnesium Hydroxide ($Mg(OH)_2$) – H2Mg1O2(g)

MW = 58.32 g/mol – $\Delta_f H^o$(298.15 K) = -572.371 kJ/mol – S^o(298.15 K) = 267.27 J/mol-K – p^o = 0.1 MPa



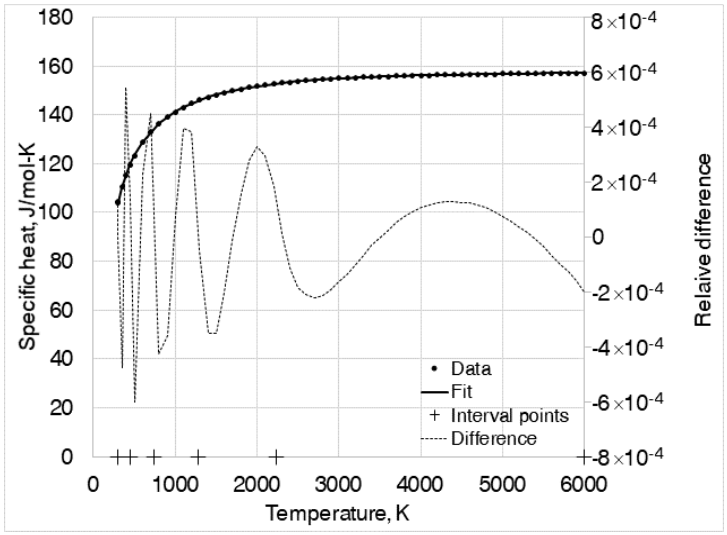


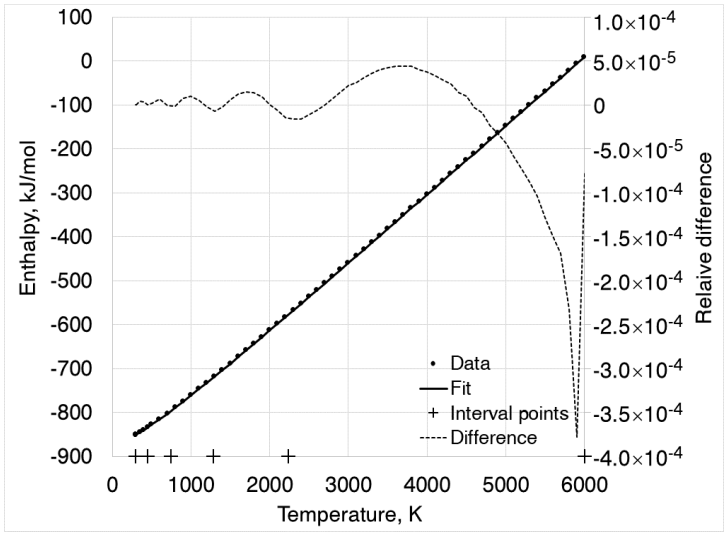


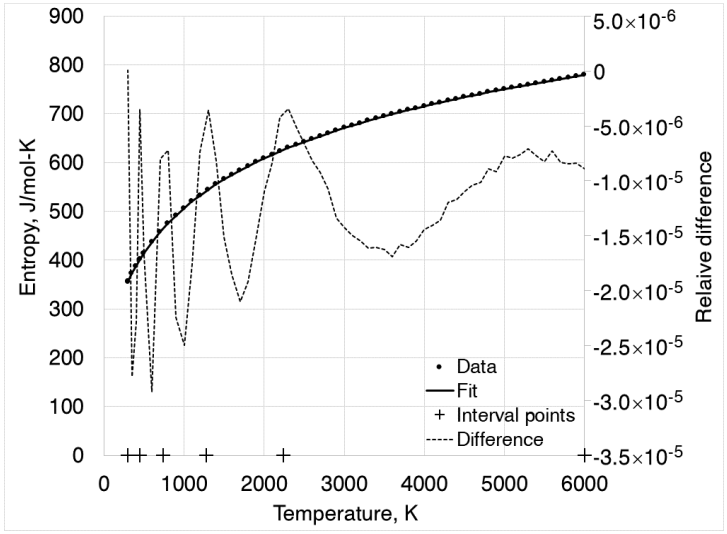


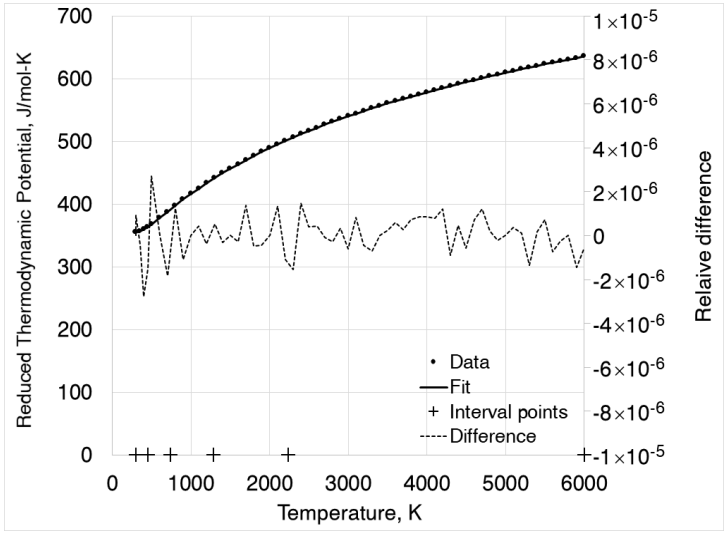
Molybdic Acid ($O_2Mo(OH)_2$) – H2Mo1O4(g)

MW = 161.95 g/mol – $\Delta_f H^o$(298.15 K) = -851.026 kJ/mol – S^o(298.15 K) = 355.62 J/mol-K – p^o = 0.1 MPa



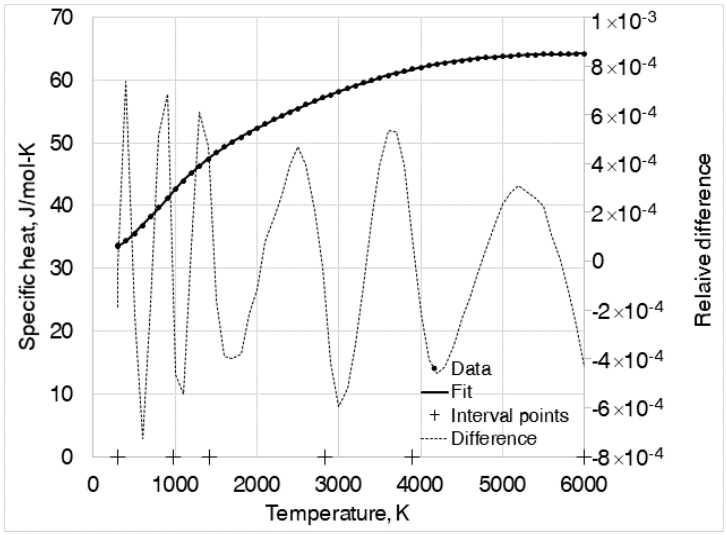


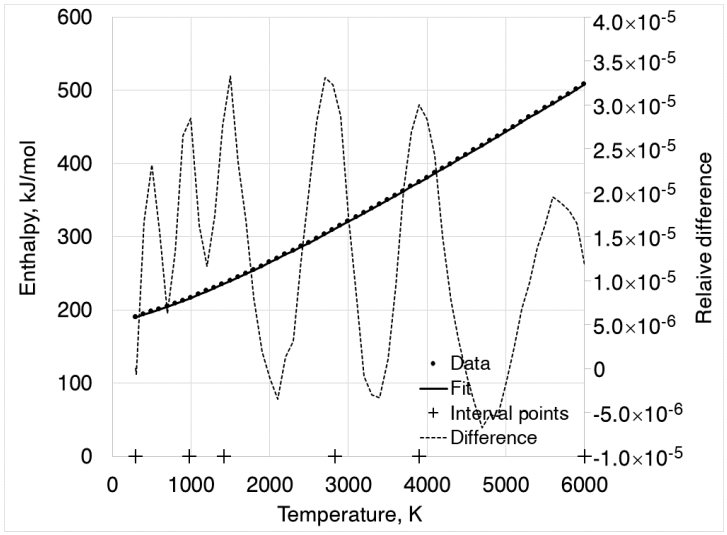


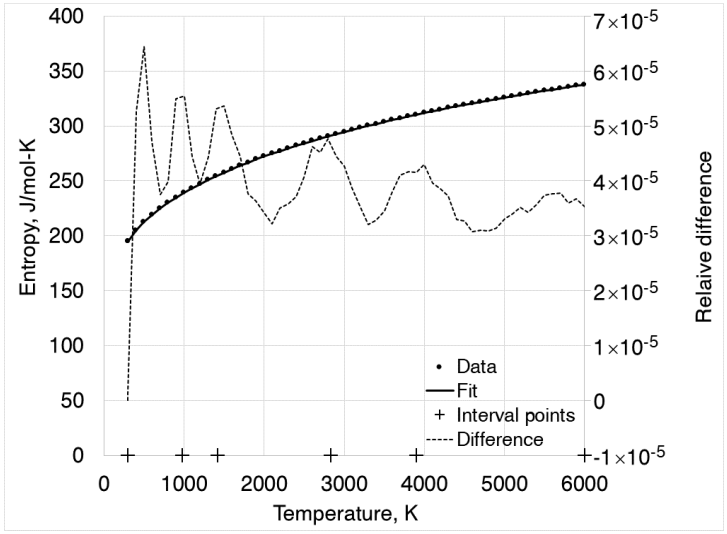


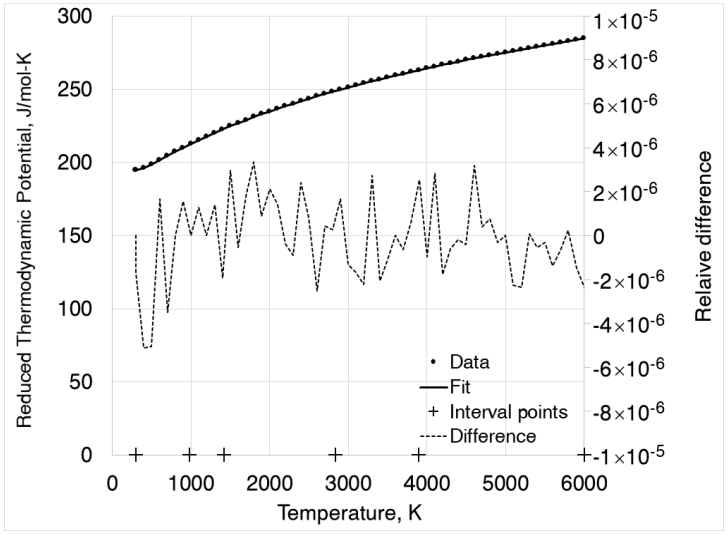
Amidogen (NH_2) – H2N1(g)

MW = 16.0226 g/mol – $\Delta_f H^o$(298.15 K) = 190.372 kJ/mol – S^o(298.15 K) = 194.707 J/mol-K – p^o = 0.1 MPa



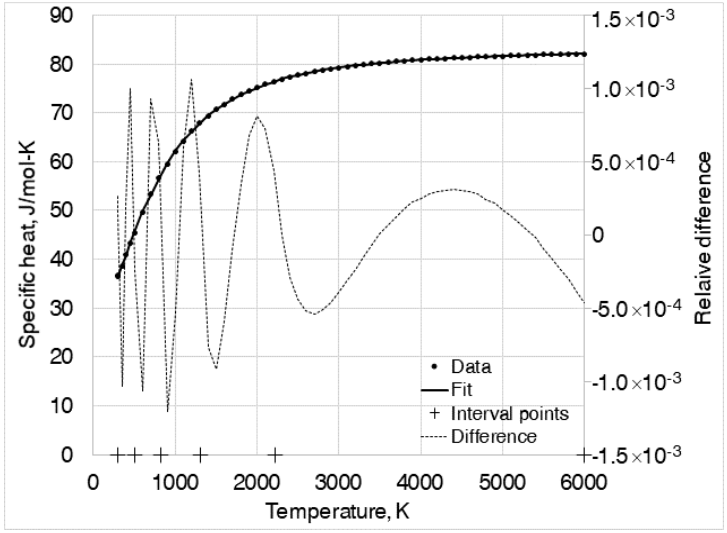


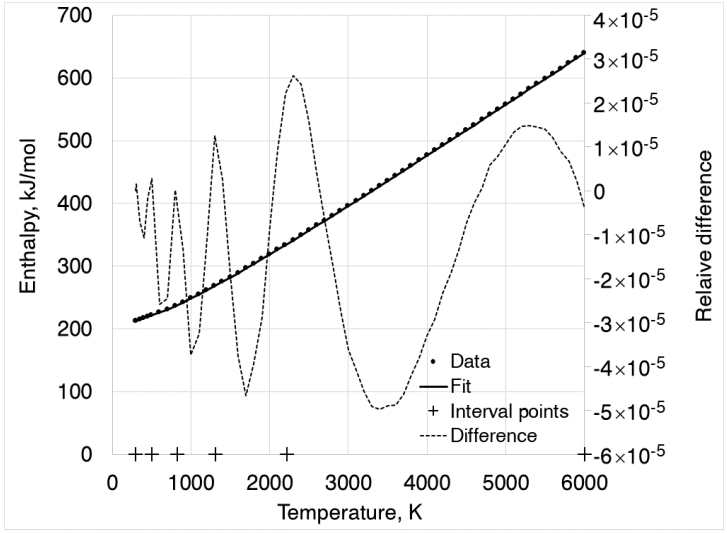


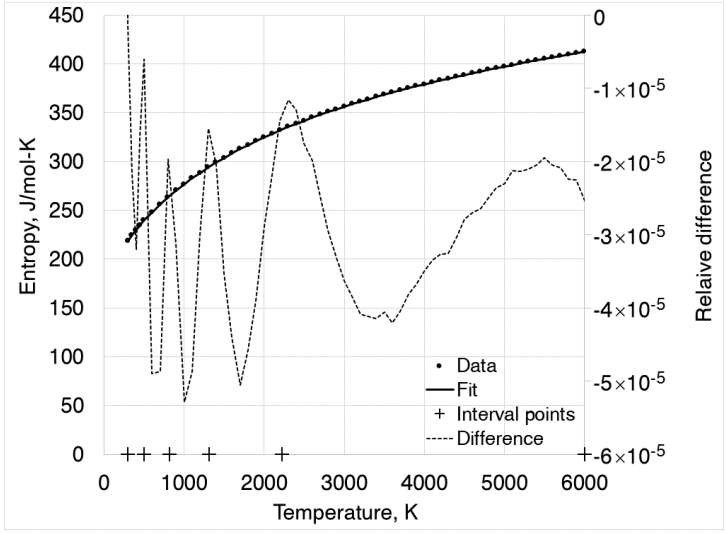


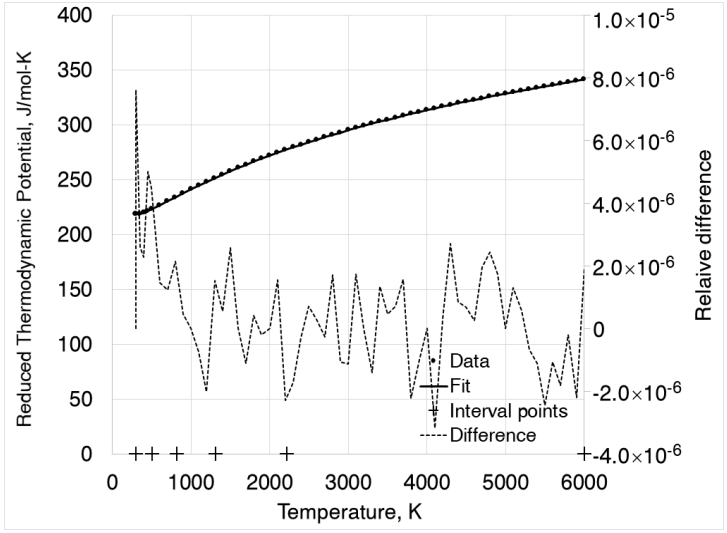
Diazene, Cis (HNNH) – H2N2(g)

MW = 30.0293 g/mol – $\Delta_f H^o$(298.15 K) = 212.966 kJ/mol – S^o(298.15 K) = 218.603 J/mol-K – p^o = 0.1 MPa



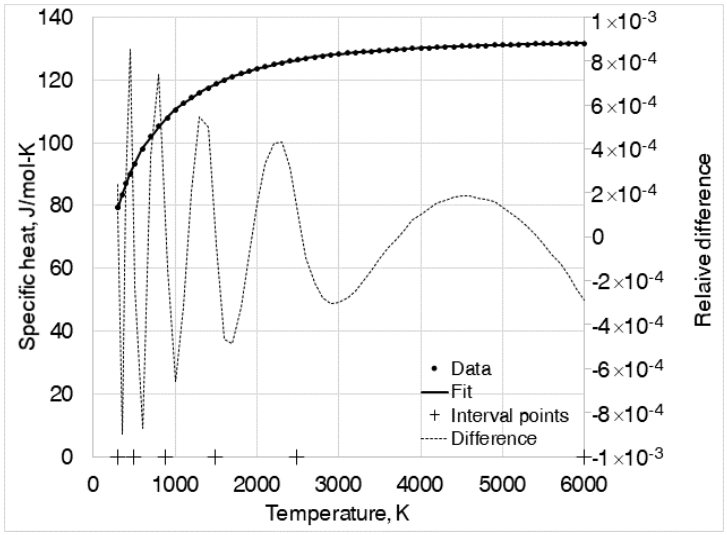


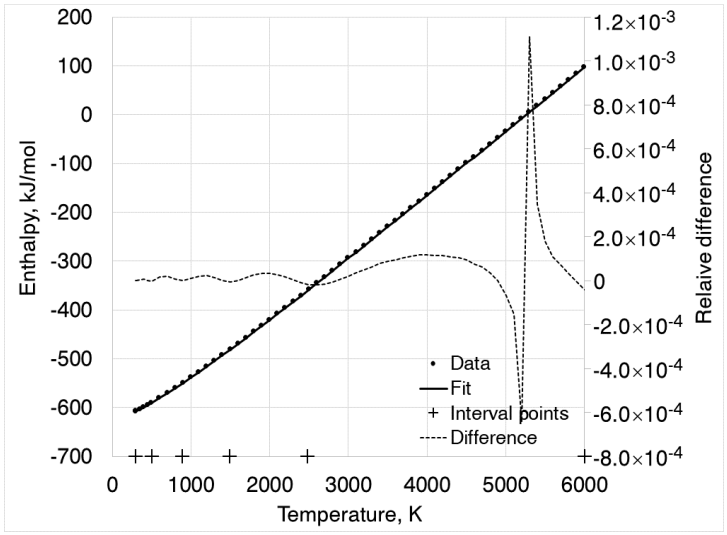


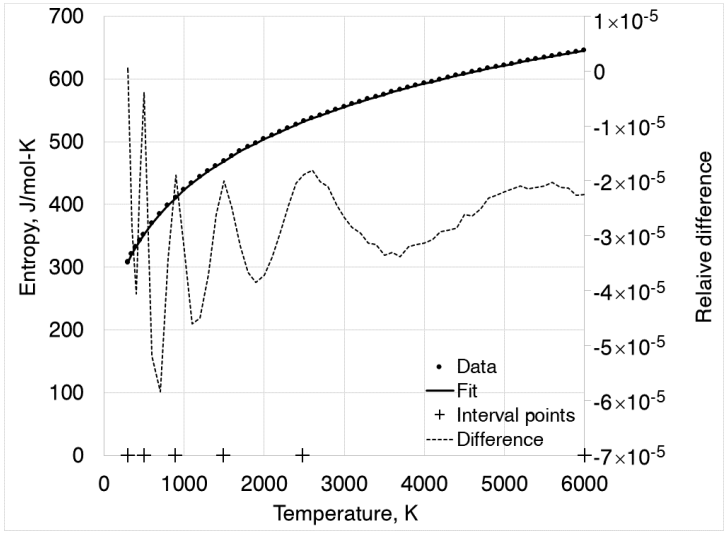


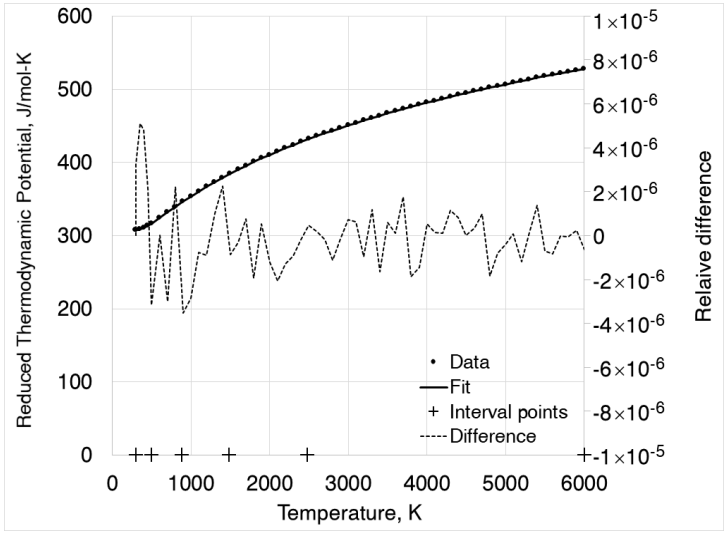
Sodium Hydroxide ((NaOH)$_2$) – H2Na2O2(g)

MW = 79.9942 g/mol – $\Delta_f H^o$(298.15 K) = -607.517 kJ/mol – S^o(298.15 K) = 307.396 J/mol-K – p^o = 0.1 MPa



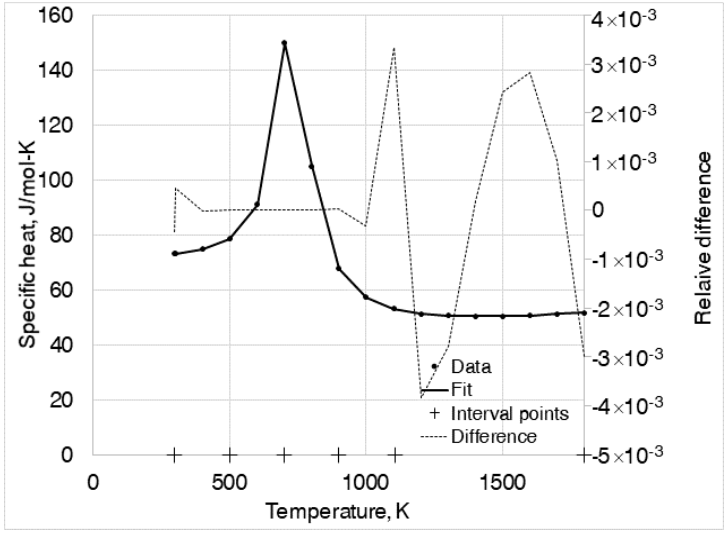


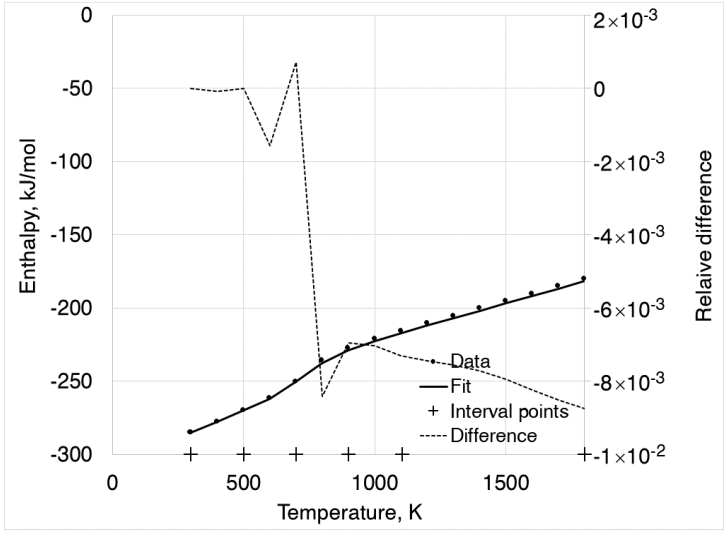


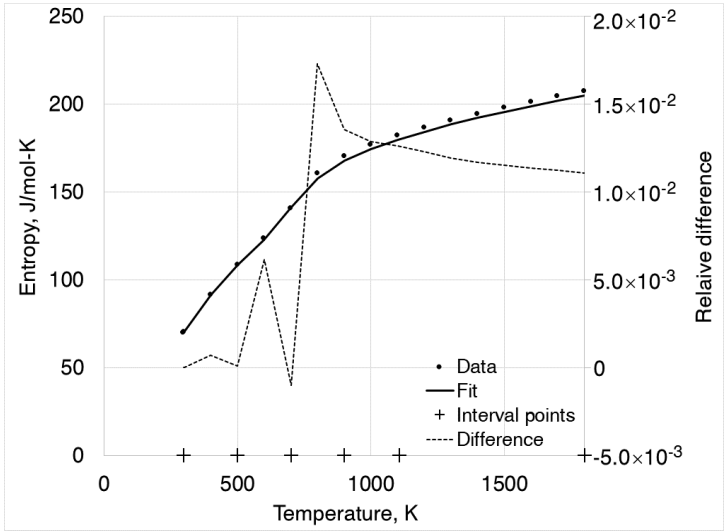


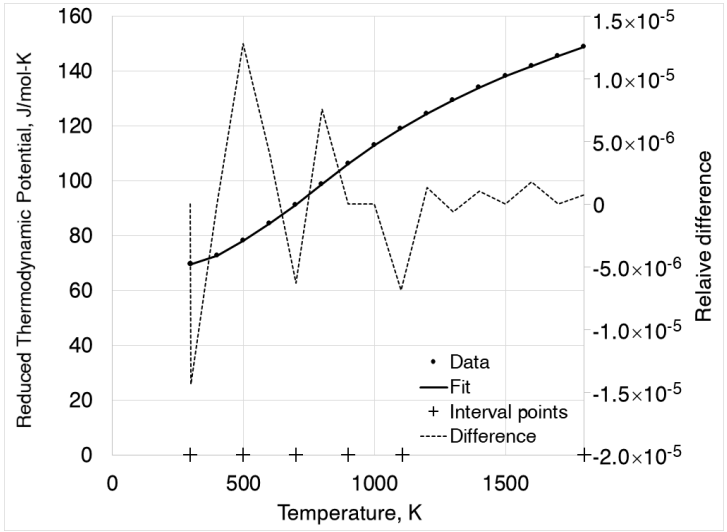
Water, 500 Bar (H_2O) – H2O1(fl)_01

MW = 18.0153 g/mol – $\Delta_f H^o$(298.15 K) = -285.014 kJ/mol – S^o(298.15 K) = 69.696 J/mol-K – p^o = 0.1 MPa



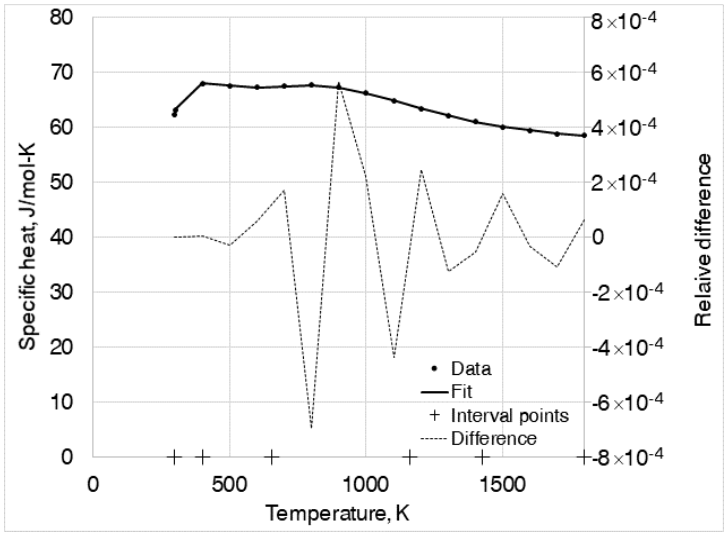


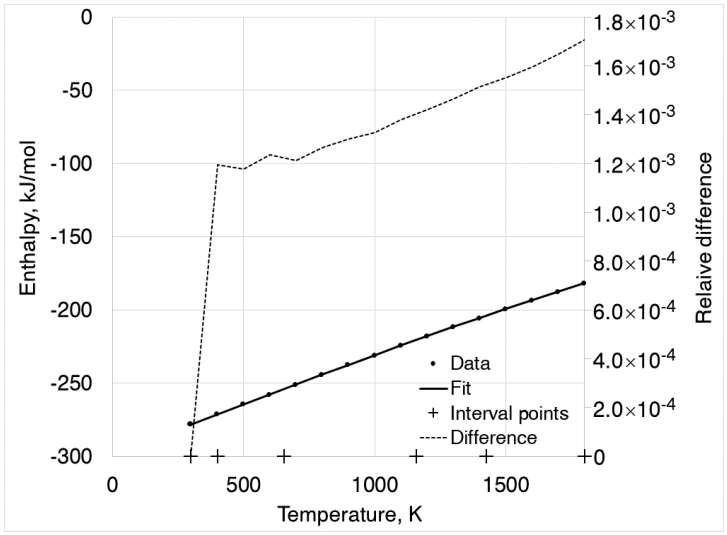


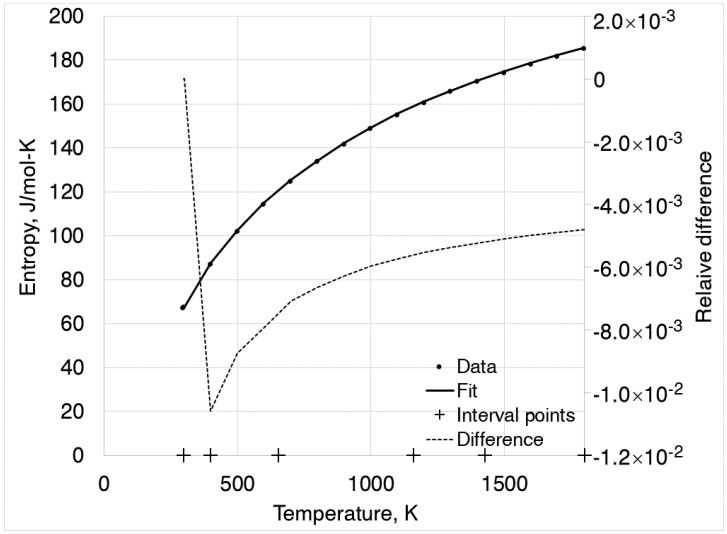


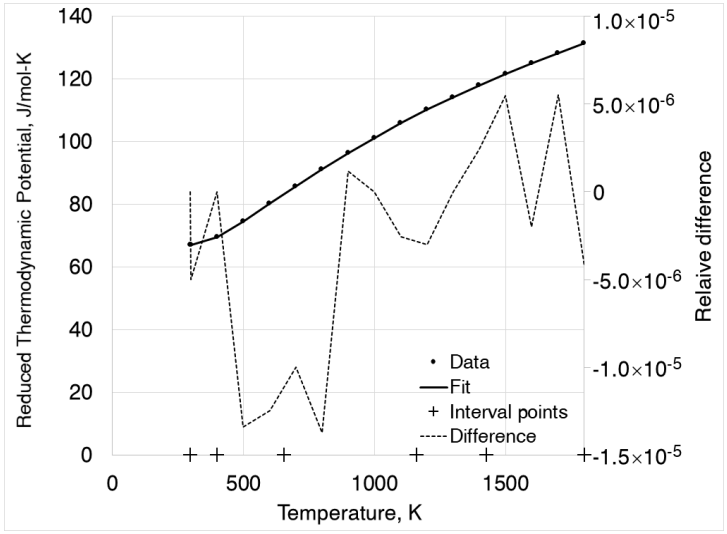
Water, 5000 Bar (H_2O) – H2O1(fl)_02

MW = 18.0153 g/mol – $\Delta_f H^o$(298.15 K) = -278.394 kJ/mol – S^o(298.15 K) = 66.927 J/mol-K – p^o = 0.1 MPa



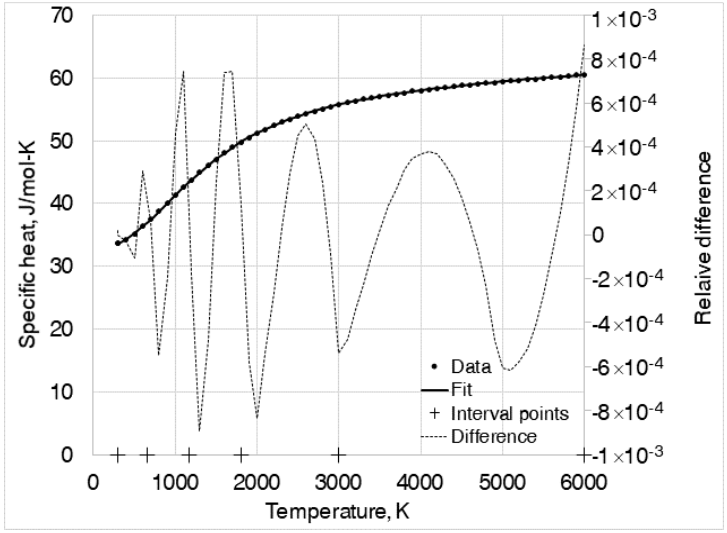


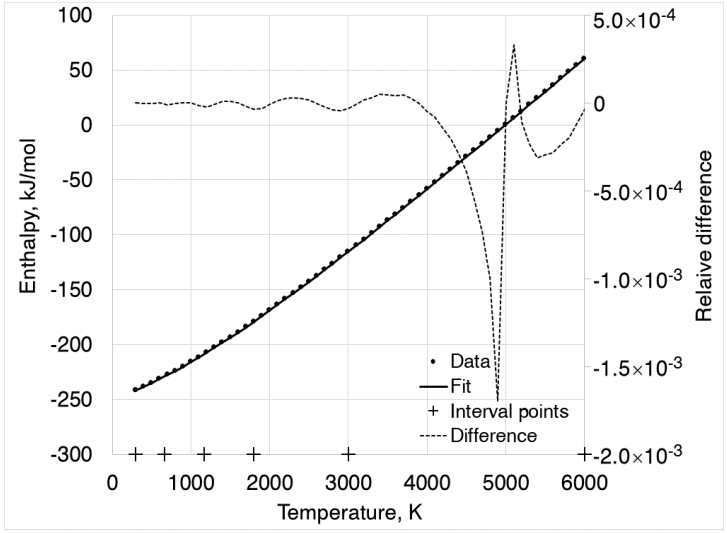


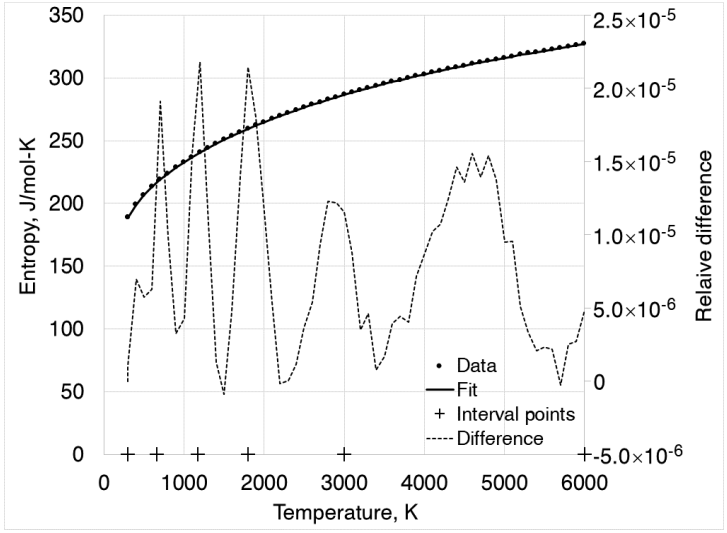


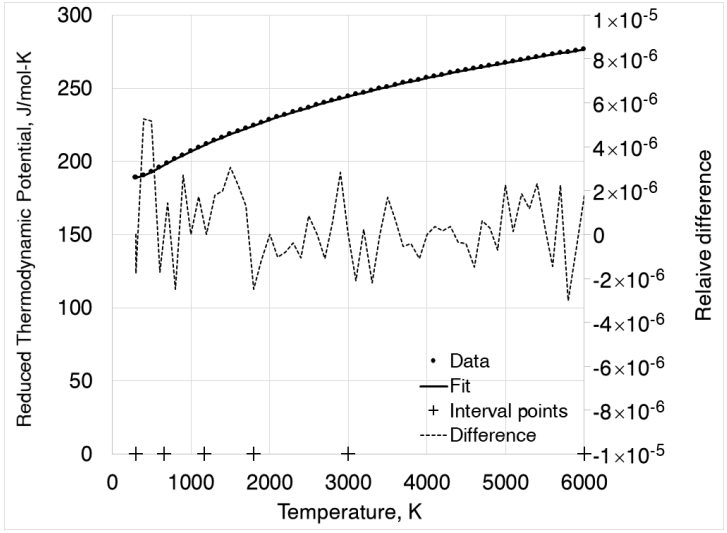
Water (H_2O) – H2O1(g)

MW = 18.0153 g/mol – $\Delta_f H^o$(298.15 K) = -241.826 kJ/mol – S^o(298.15 K) = 188.834 J/mol-K – p^o = 0.1 MPa



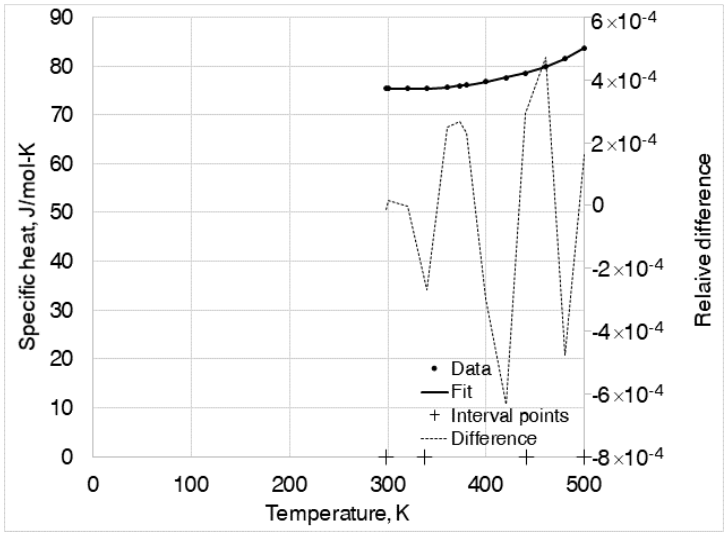


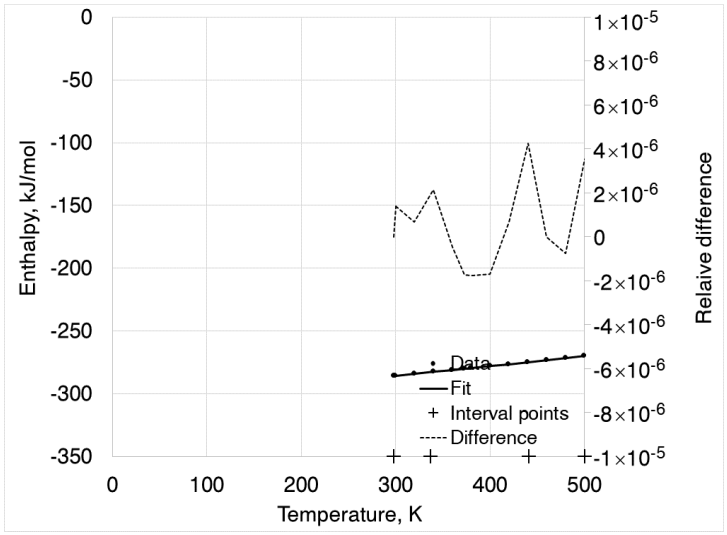


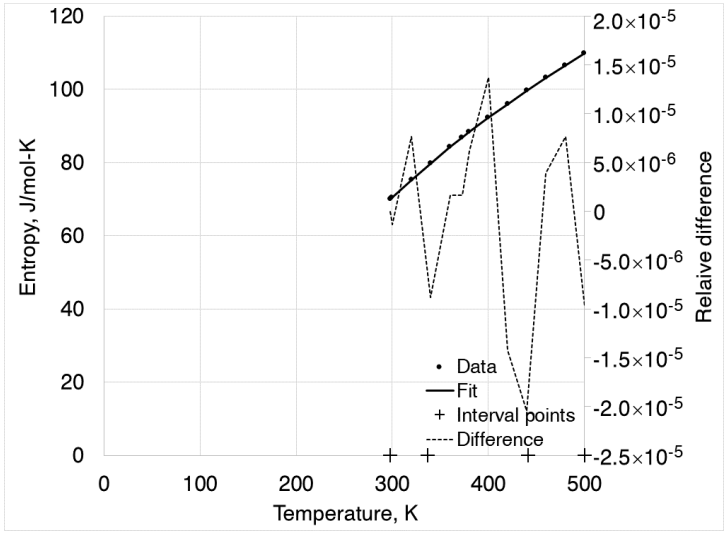


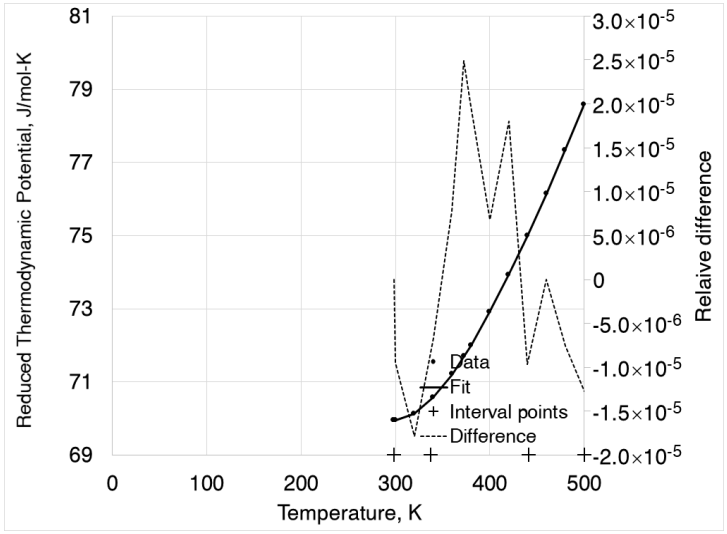
Water (H_2O) – H2O1(l)

MW = 18.0153 g/mol – $\Delta_f H^o$(298.15 K) = -285.83 kJ/mol – S^o(298.15 K) = 69.95 J/mol-K – p^o = 0.1 MPa



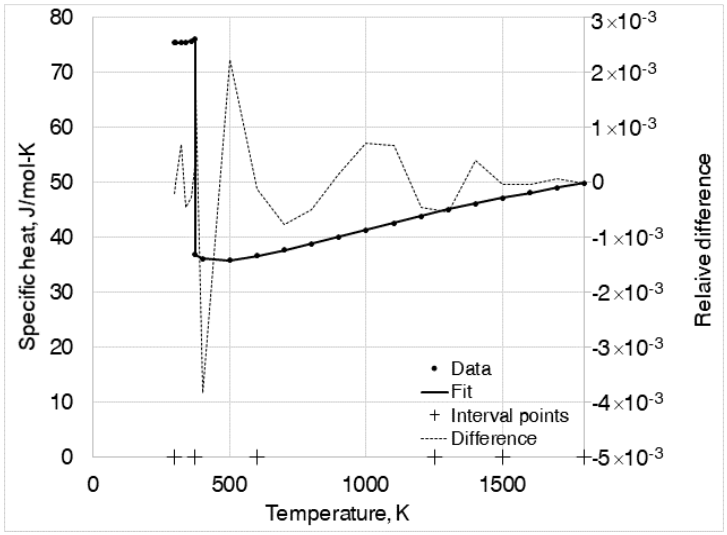


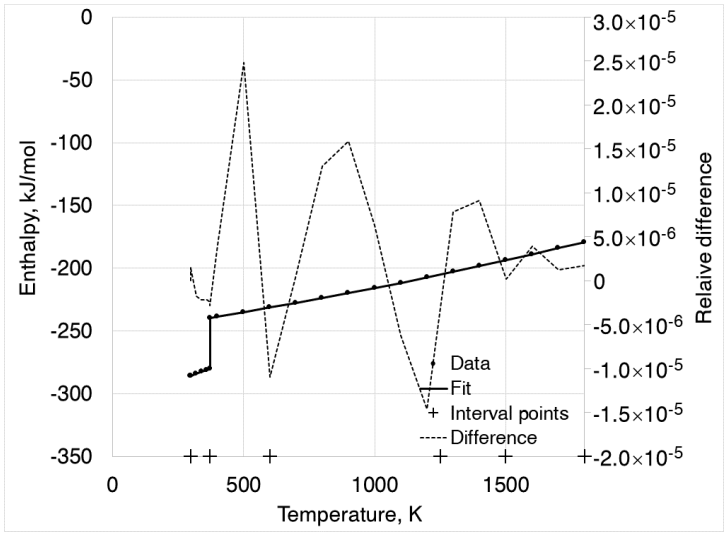


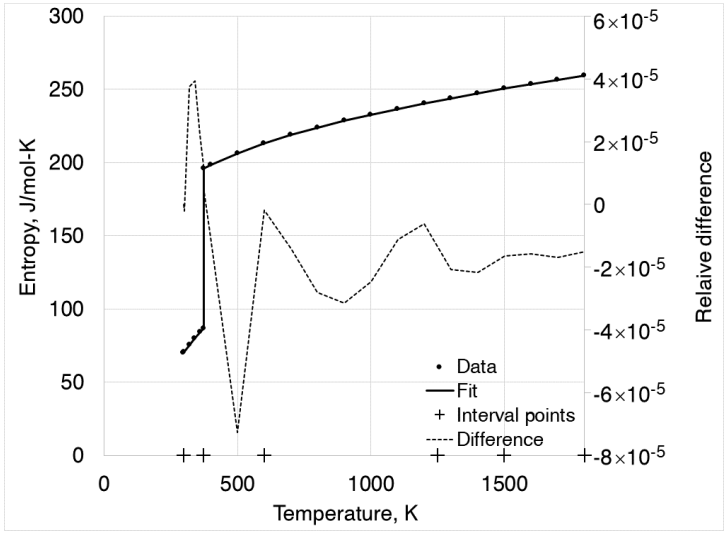


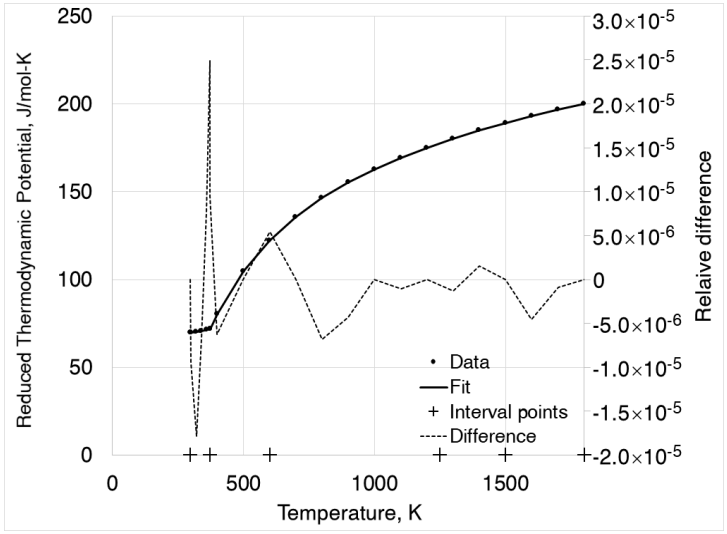
Water, 1 Bar (H_2O) – H2O1(l,g)_01

MW = 18.0153 g/mol – $\Delta_f H^o$(298.15 K) = -285.83 kJ/mol – S^o(298.15 K) = 69.95 J/mol-K – p^o = 0.1 MPa



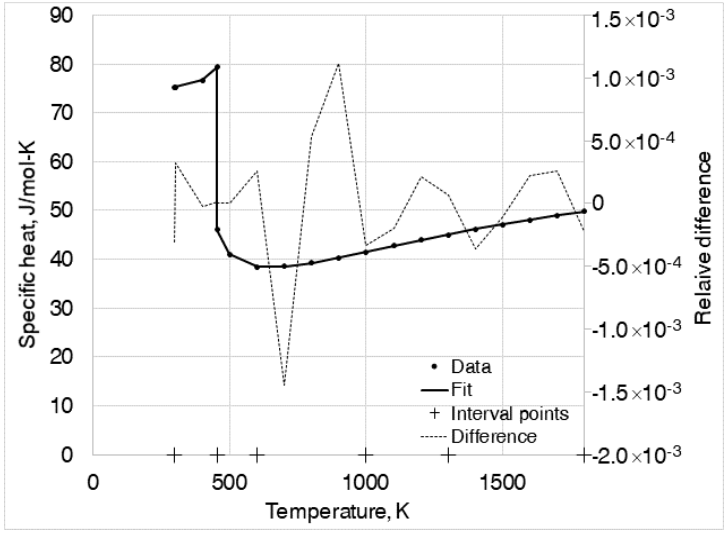


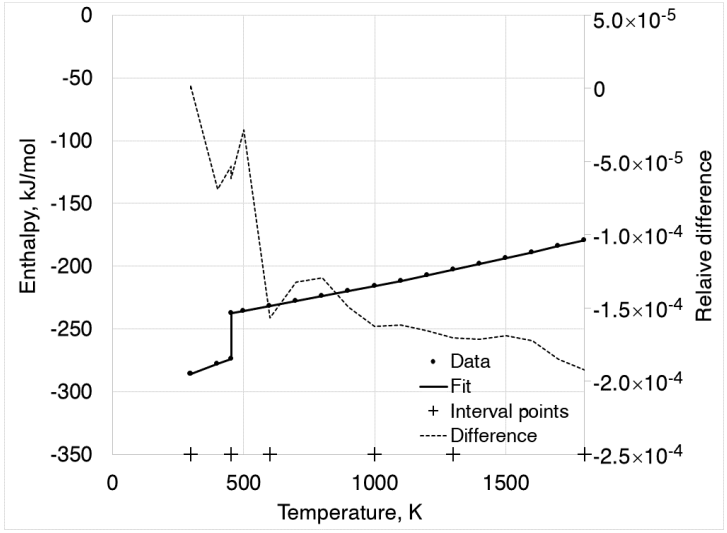


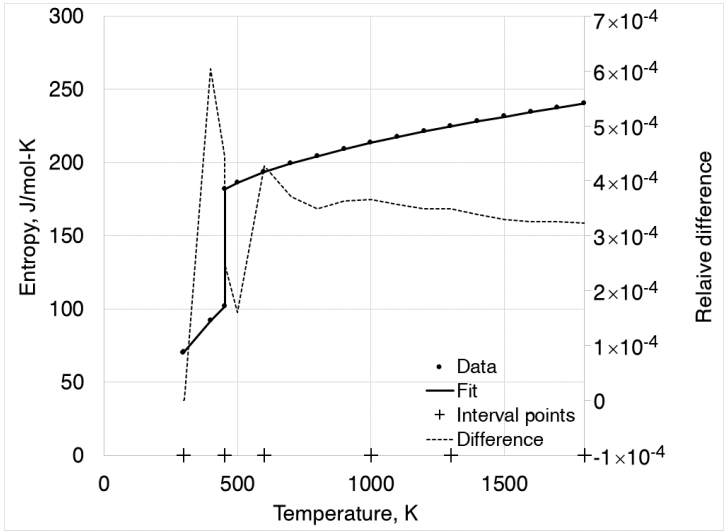


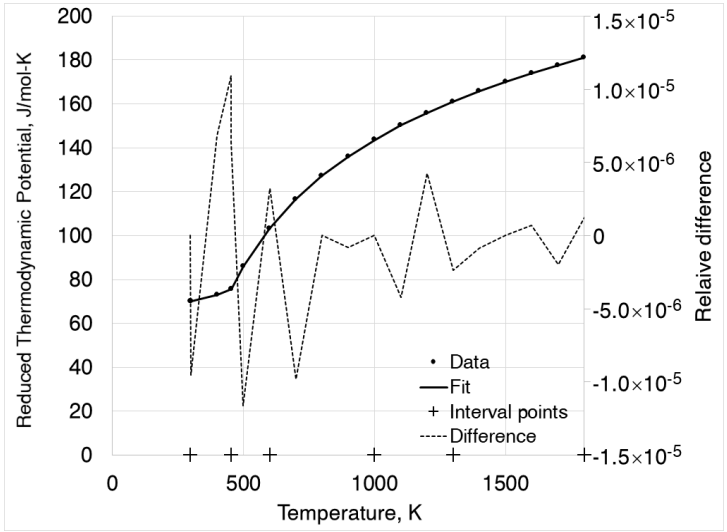
Water, 10 Bar (H_2O) – H2O1(l,g)_02

MW = 18.0153 g/mol – $\Delta_f H^o$(298.15 K) = -285.815 kJ/mol – S^o(298.15 K) = 69.946 J/mol-K – p^o = 0.1 MPa



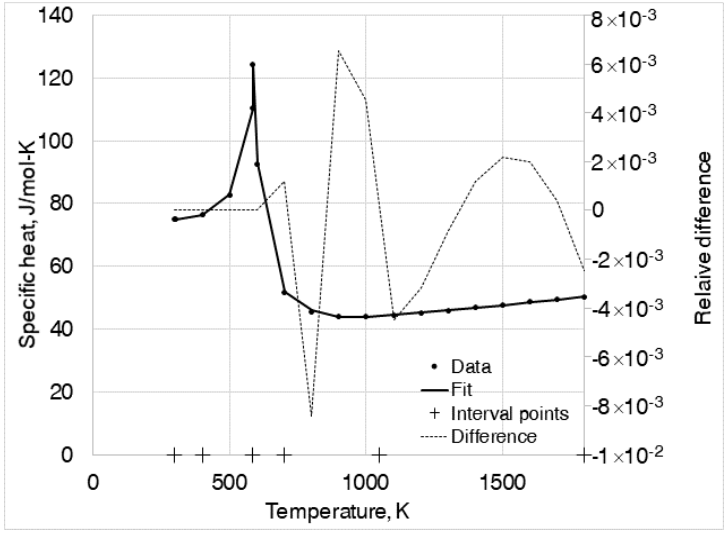


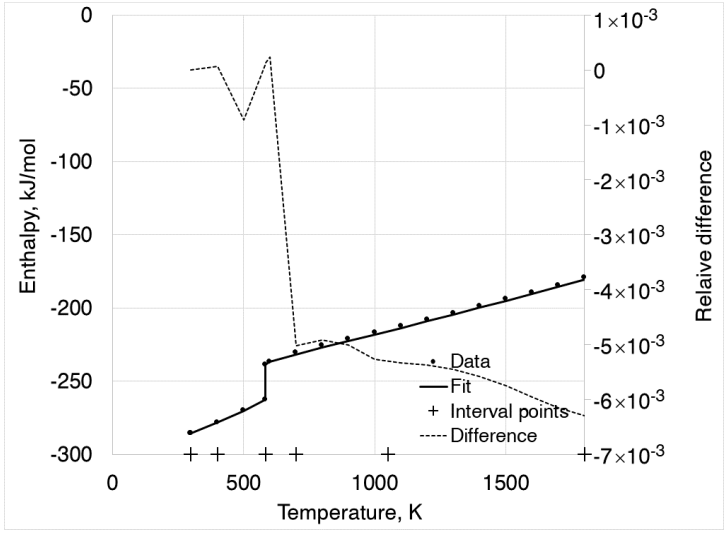


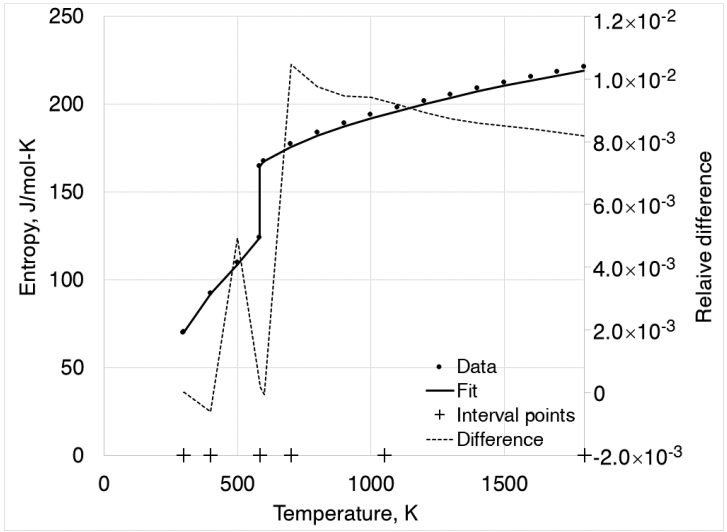


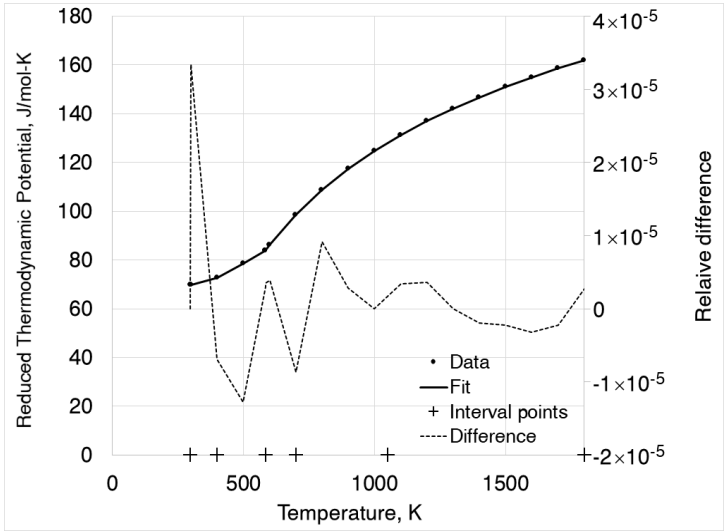
Water, 100 Bar (H_2O) – H2O1(l,g)_03

MW = 18.0153 g/mol – $\Delta_f H^o$(298.15 K) = -285.666 kJ/mol – S^o(298.15 K) = 69.903 J/mol-K – p^o = 0.1 MPa



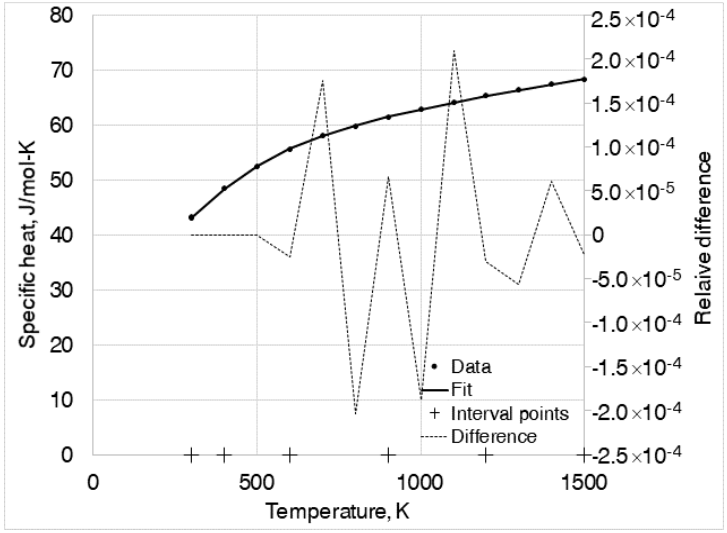


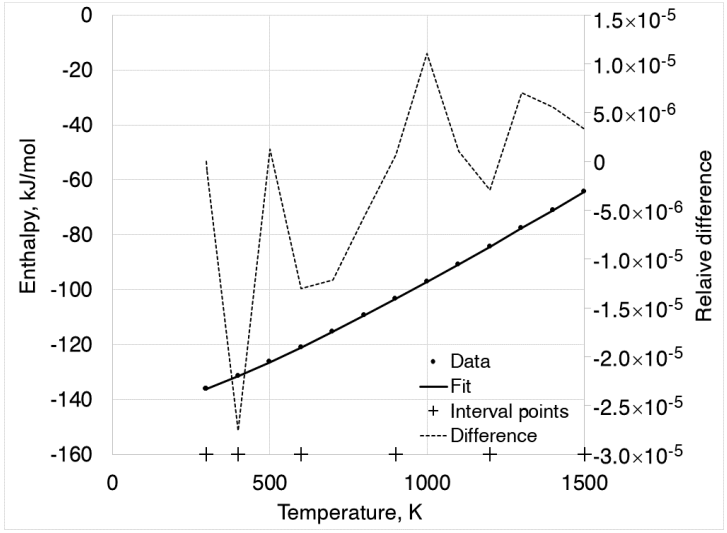


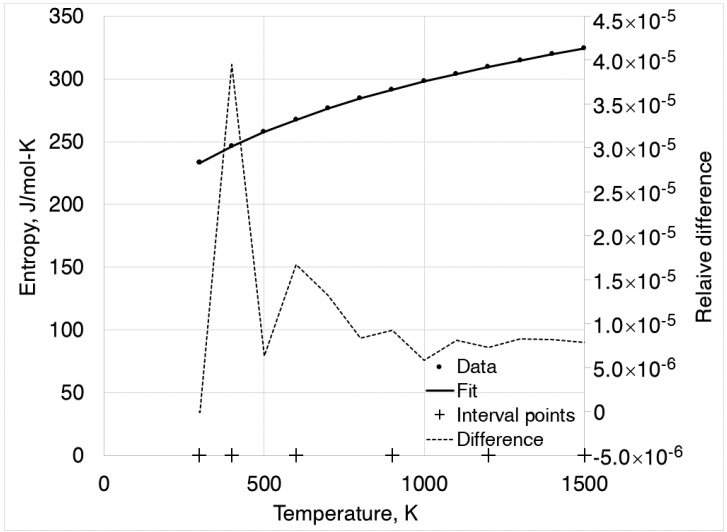


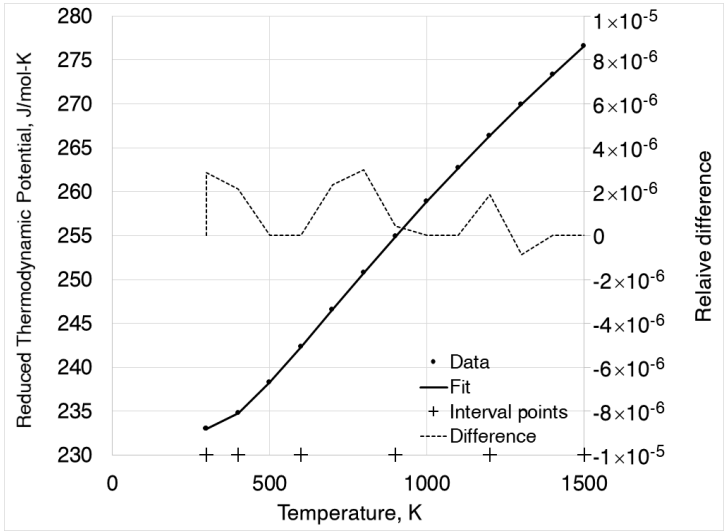
Hydrogen Peroxide (HOOH) – H2O2(g)

MW = 34.0147 g/mol – $\Delta_f H^o$(298.15 K) = -136.106 kJ/mol – S^o(298.15 K) = 232.991 J/mol-K – p^o = 0.1 MPa



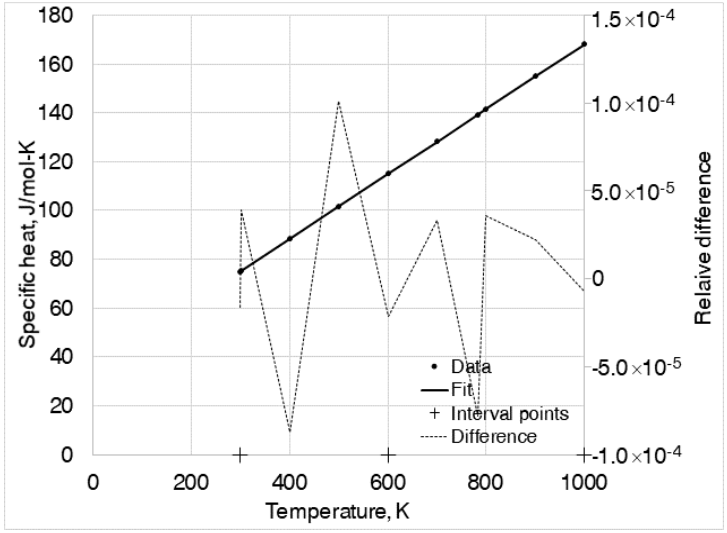


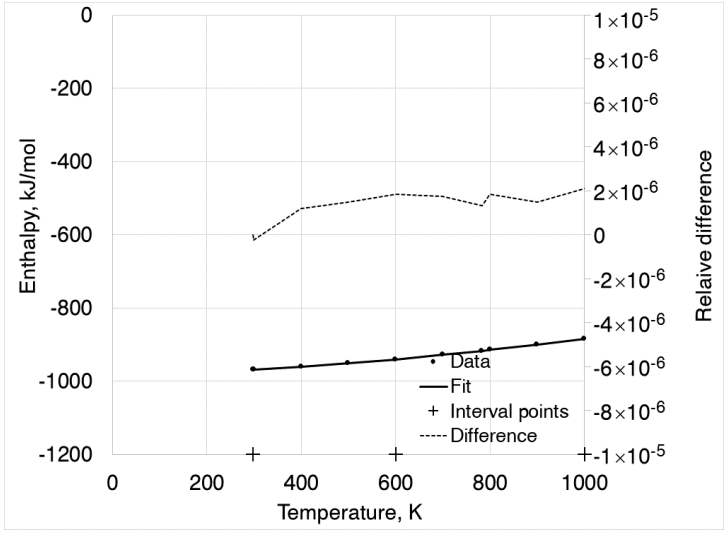


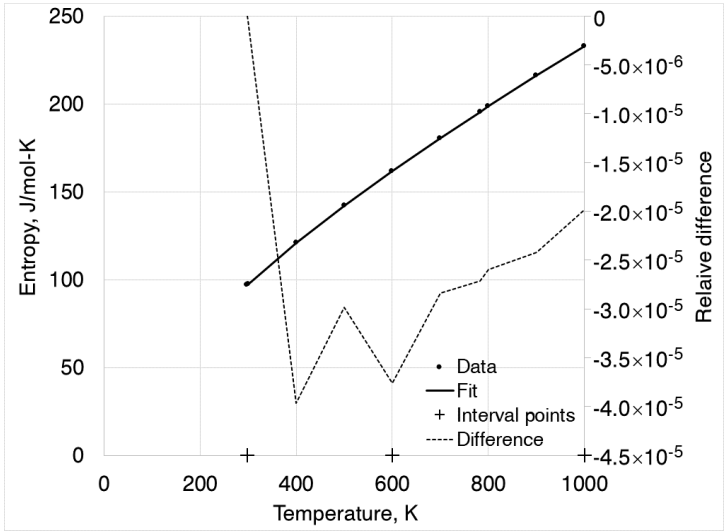


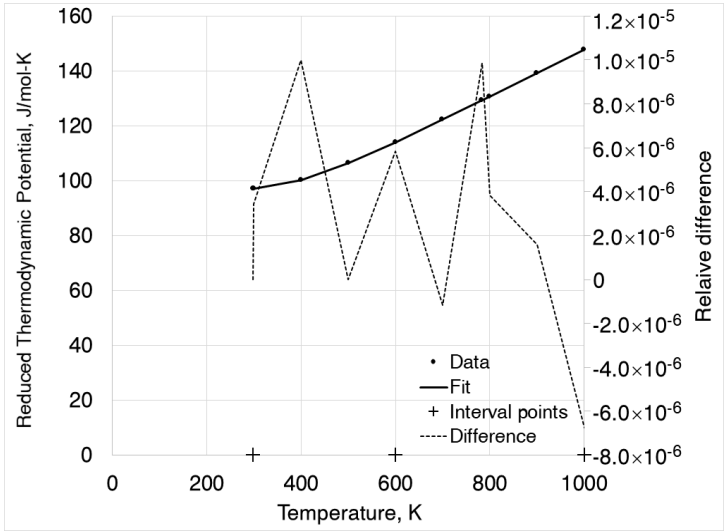
Strontium Hydroxide ($Sr(OH)_2$) – H2O2Sr1(cr)

MW = 121.63 g/mol – $\Delta_f H^o$(298.15 K) = -968.889 kJ/mol – S^o(298.15 K) = 97.069 J/mol-K – p^o = 0.1 MPa



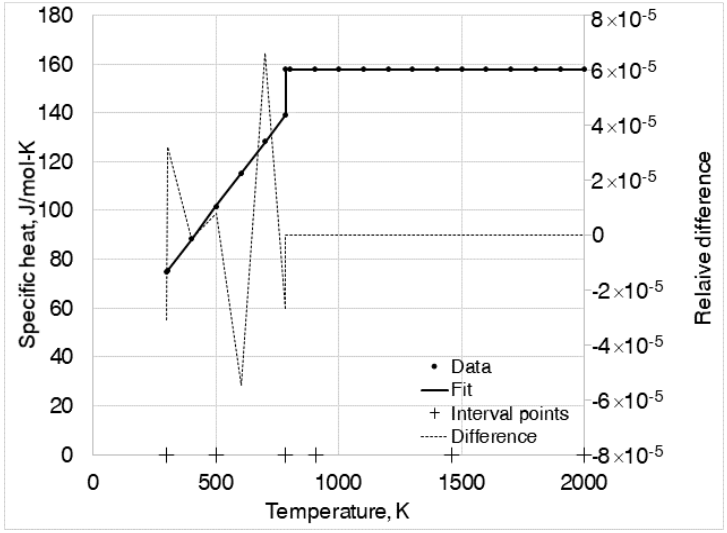


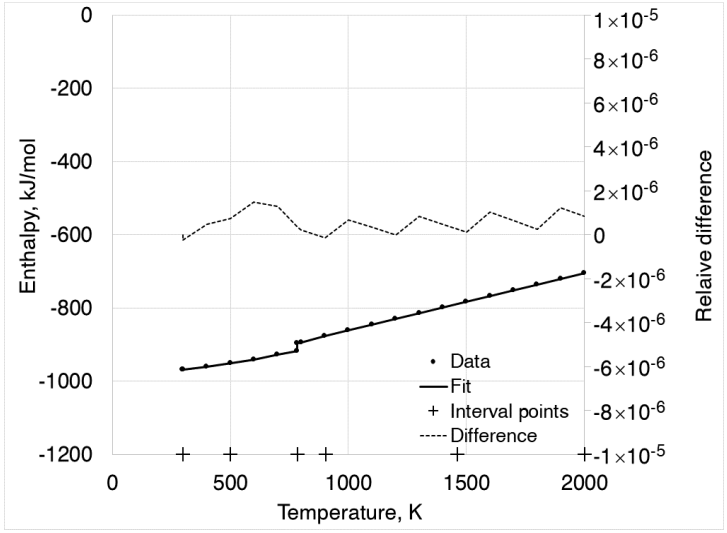


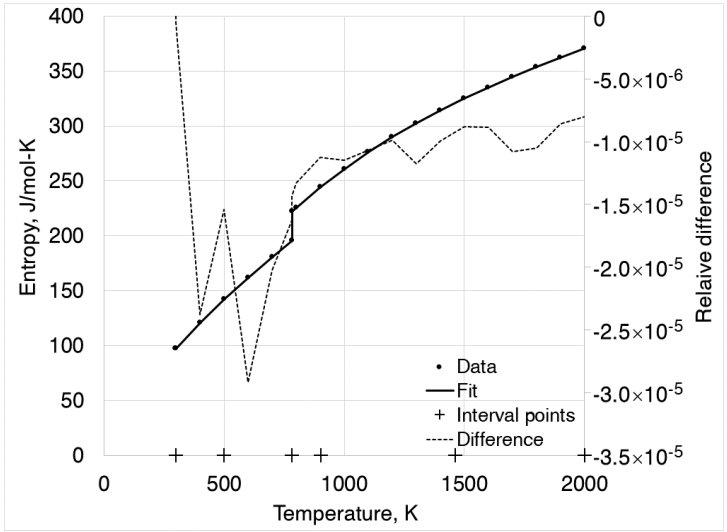


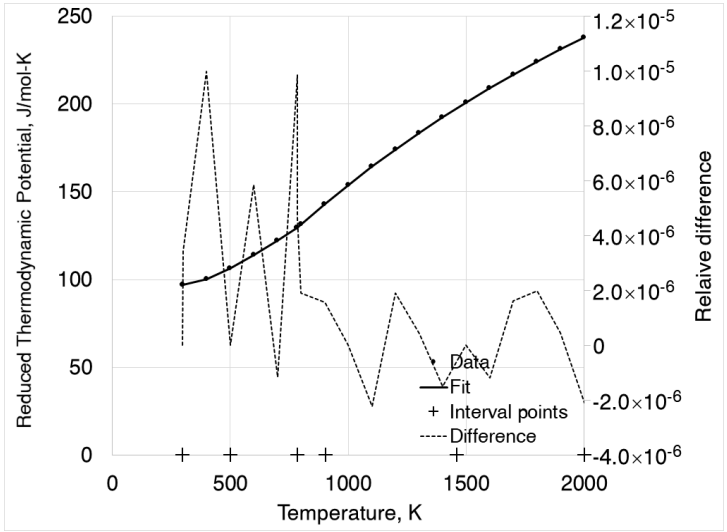
Strontium Hydroxide ($Sr(OH)_2$) – H2O2Sr1(cr,l)

MW = 121.63 g/mol – $\Delta_f H^o$(298.15 K) = -968.889 kJ/mol – S^o(298.15 K) = 97.069 J/mol-K – p^o = 0.1 MPa



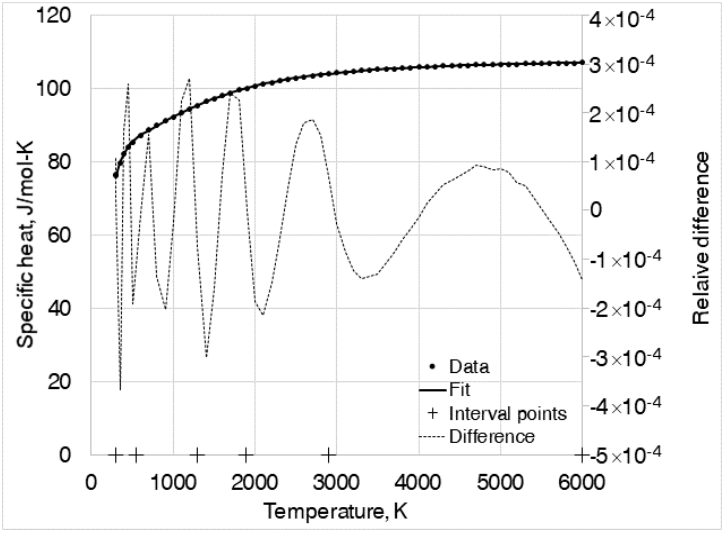


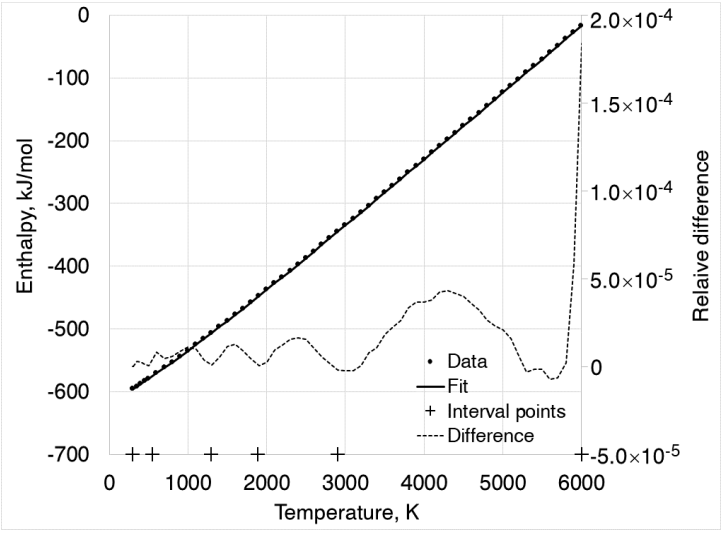


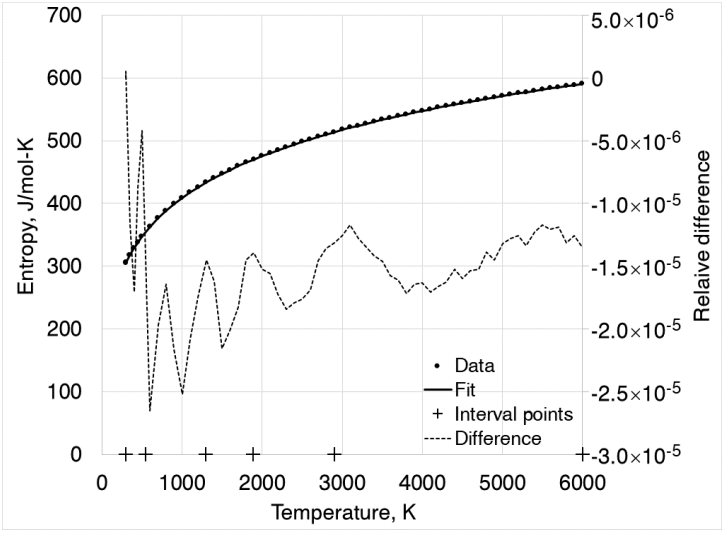


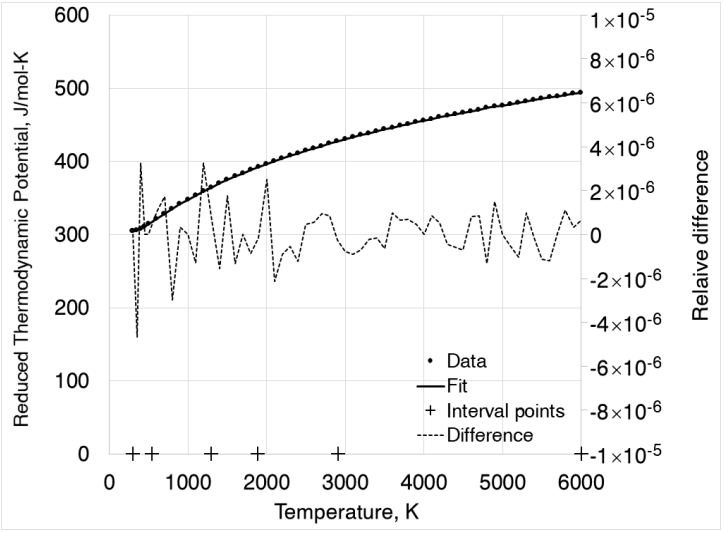
Strontium Hydroxide ($Sr(OH)_2$) – H2O2Sr1(g)

MW = 121.63 g/mol – $\Delta_f H^o$(298.15 K) = -595.797 kJ/mol – S^o(298.15 K) = 305.042 J/mol-K – p^o = 0.1 MPa



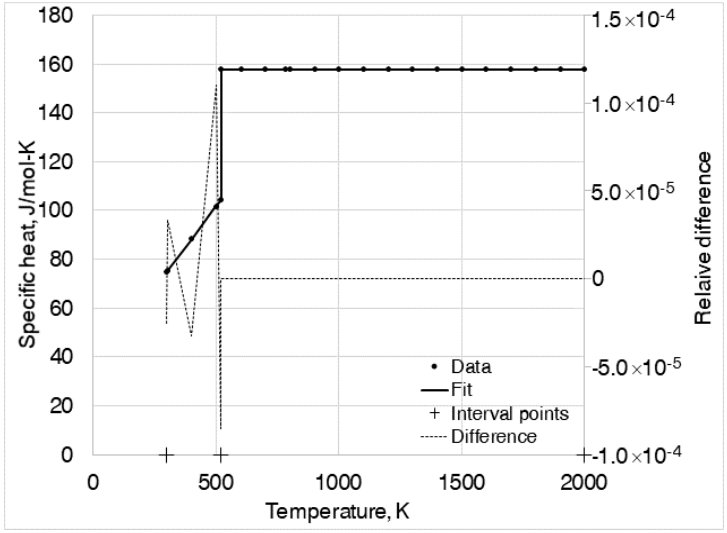


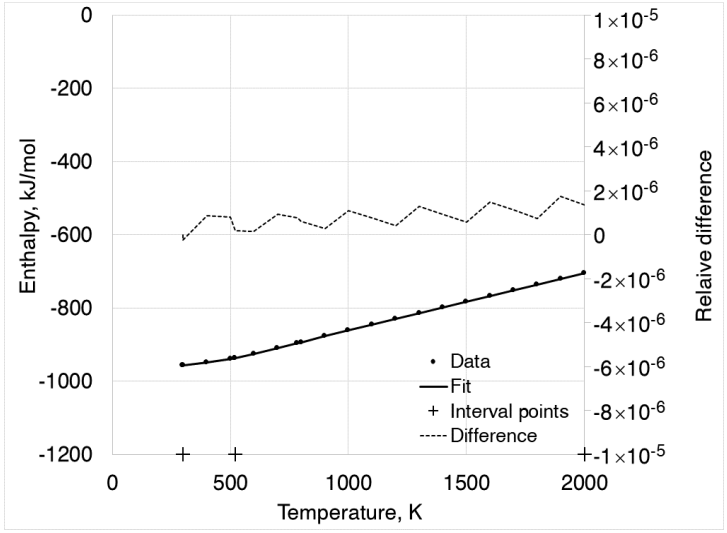


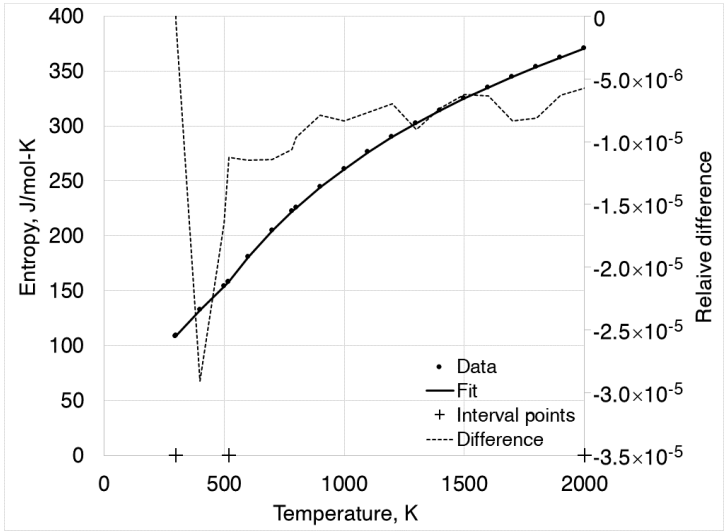


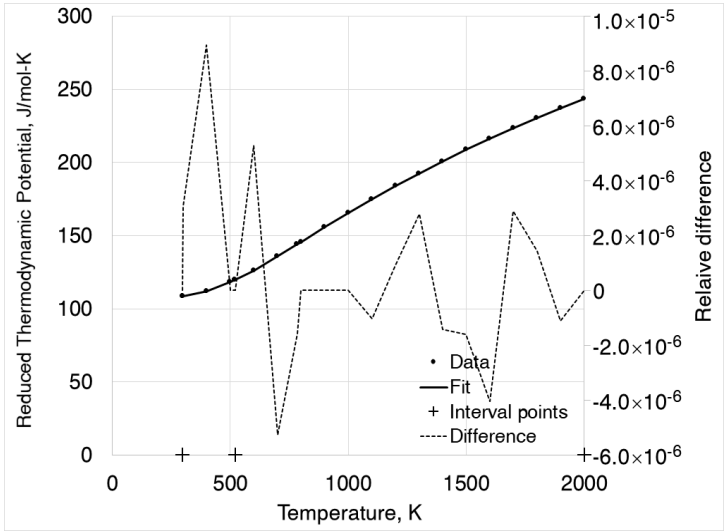
Strontium Hydroxide ($Sr(OH)_2$) – H2O2Sr1(l)

MW = 121.63 g/mol – $\Delta_f H^o$(298.15 K) = -957.304 kJ/mol – S^o(298.15 K) = 108.739 J/mol-K – p^o = 0.1 MPa



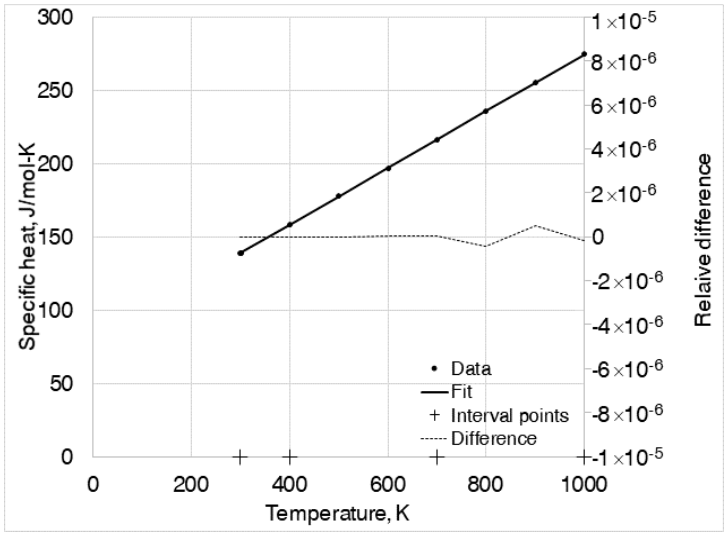


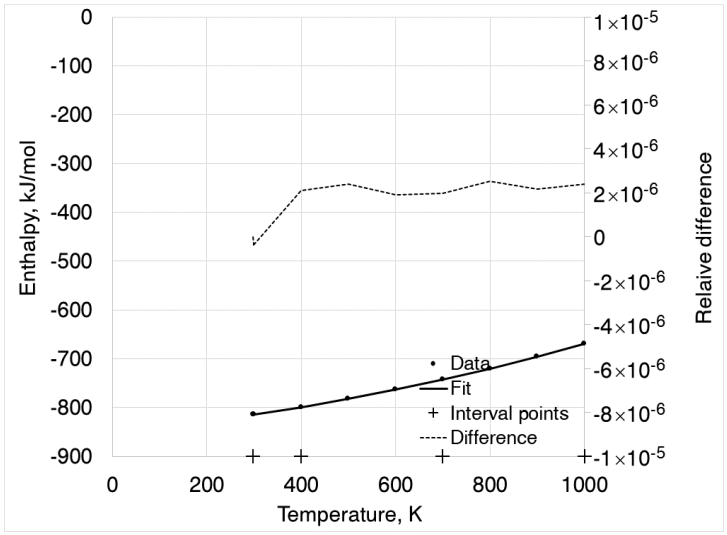


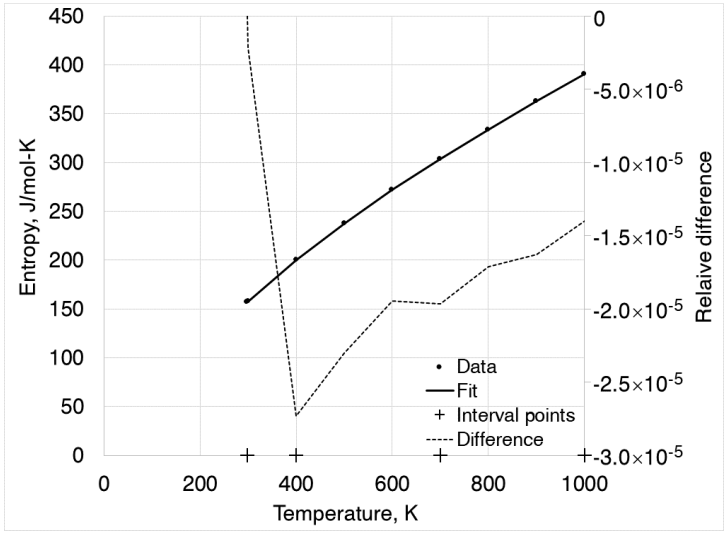


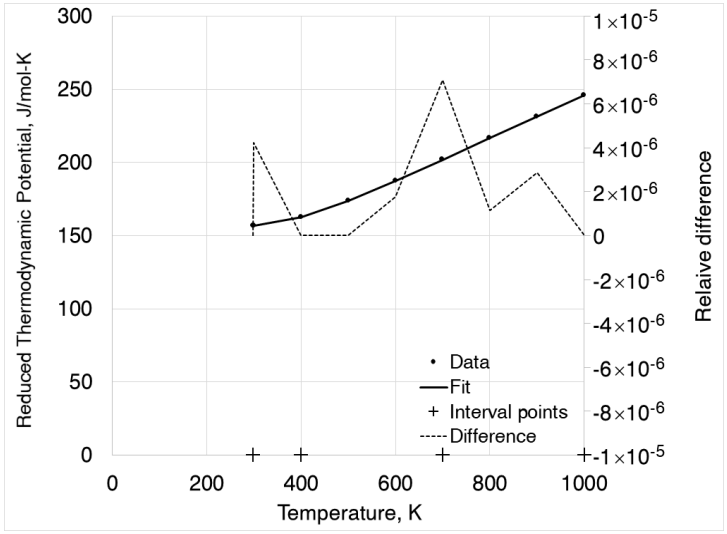
Sulfuric Acid ($O_2S(OH)_2$) – H2O4S1(cr,l)

MW = 98.07 g/mol – $\Delta_f H^o$(298.15 K) = -813.989 kJ/mol – S^o(298.15 K) = 156.895 J/mol-K – p^o = 0.1 MPa



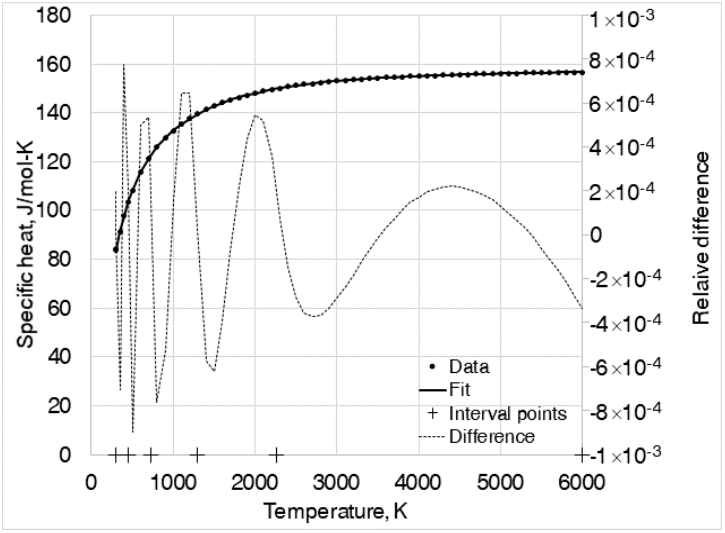


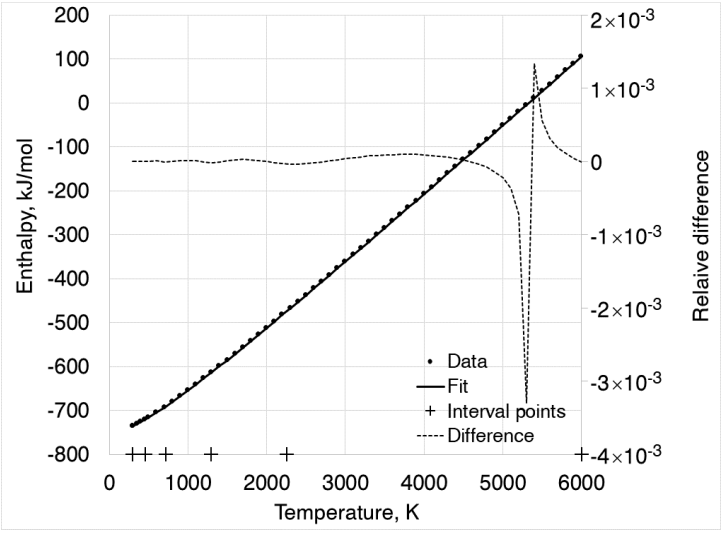


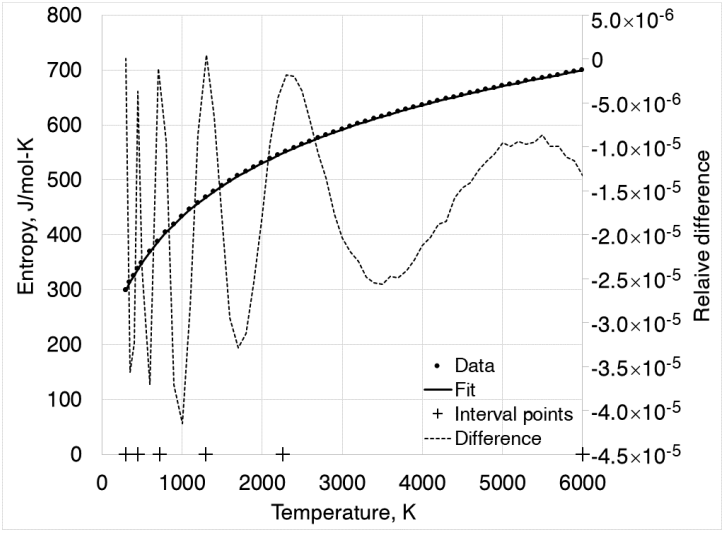


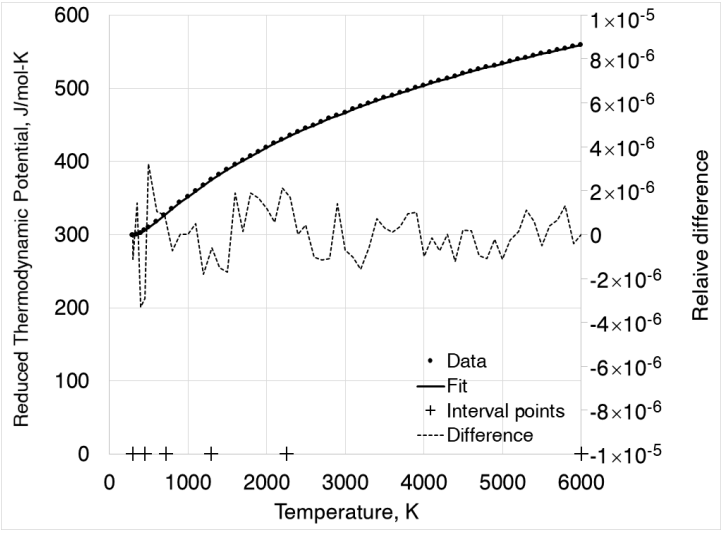
Sulfuric Acid ($O_2S(OH)_2$) – H2O4S1(g)

MW = 98.07 g/mol – $\Delta_f H^o$(298.15 K) = -735.129 kJ/mol – S^o(298.15 K) = 298.796 J/mol-K – p^o = 0.1 MPa



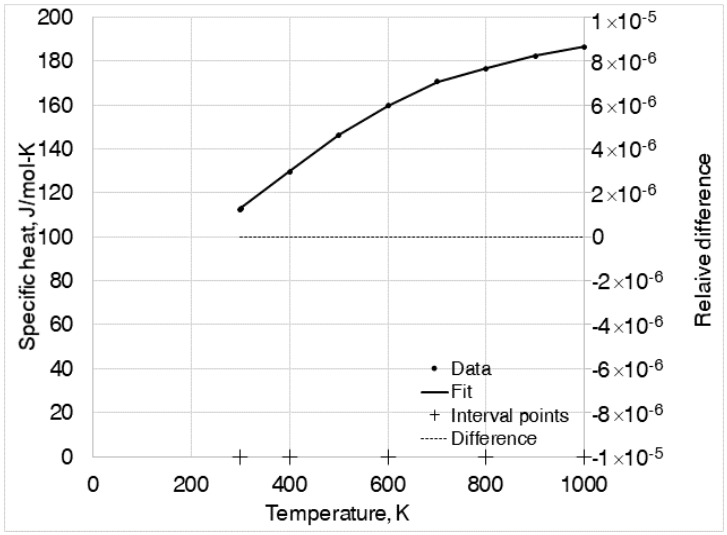


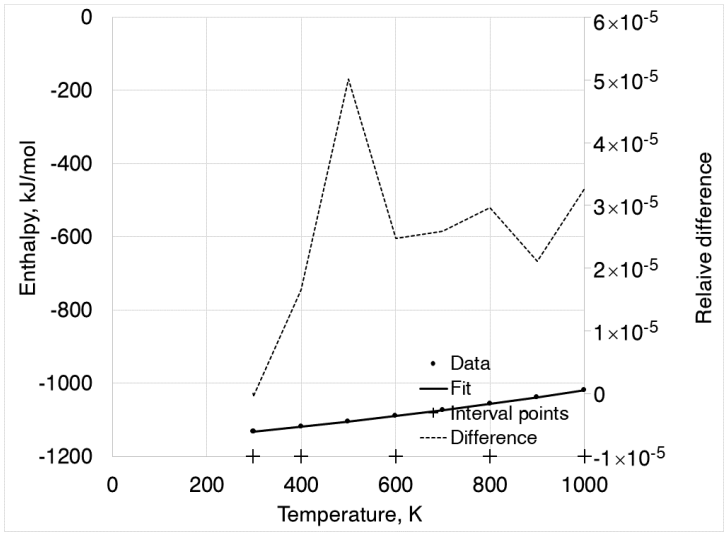


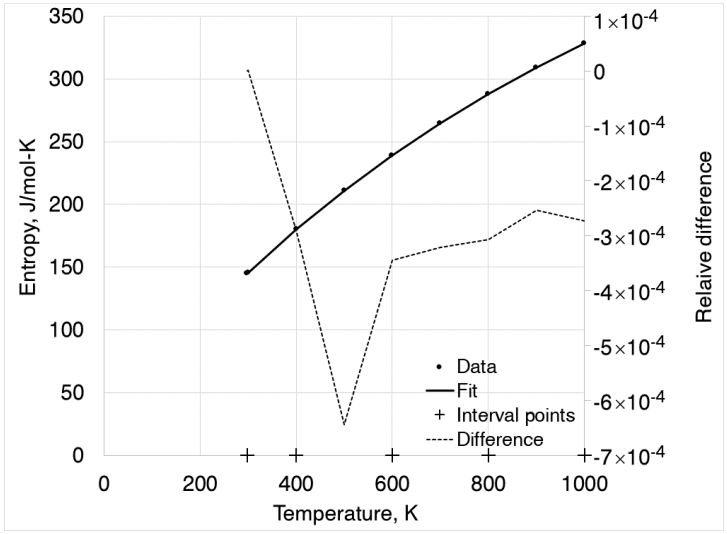


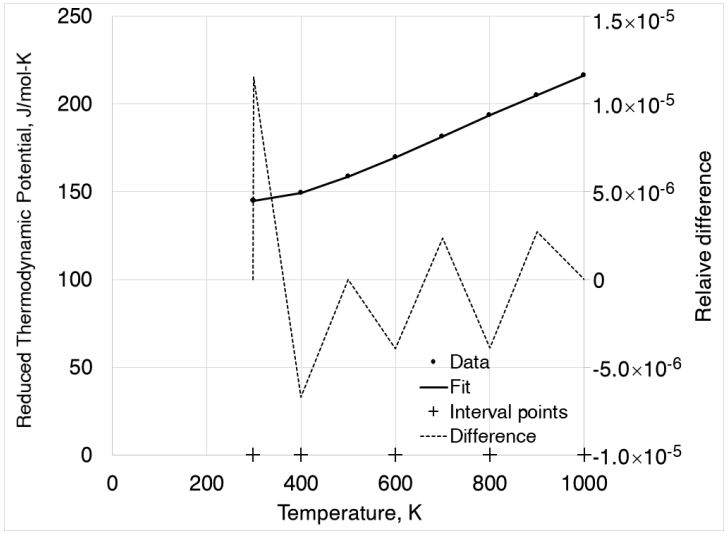
Tungstic Acid ($O_2W(OH)_2$) – H2O4W1(cr)

MW = 249.86 g/mol – $\Delta_f H^o$(298.15 K) = -1131.772 kJ/mol – S^o(298.15 K) = 144.766 J/mol-K – p^o = 0.1 MPa



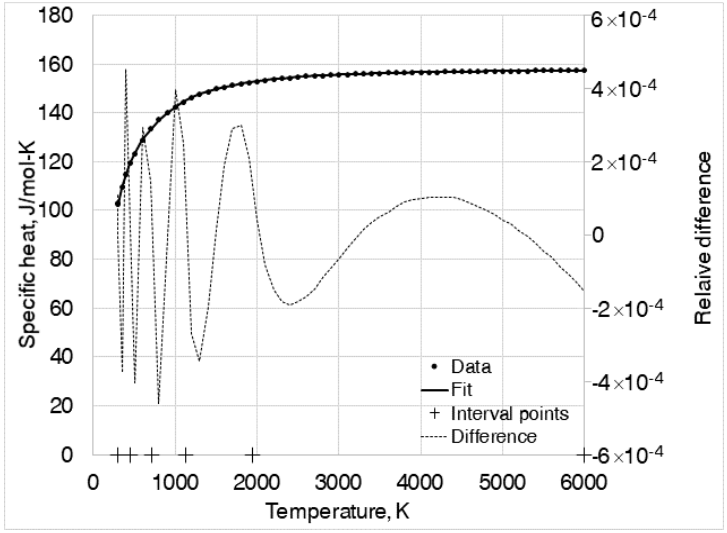


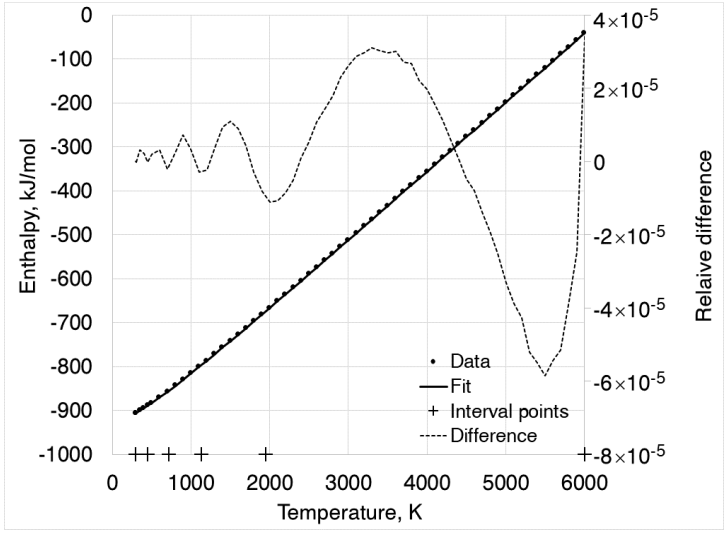


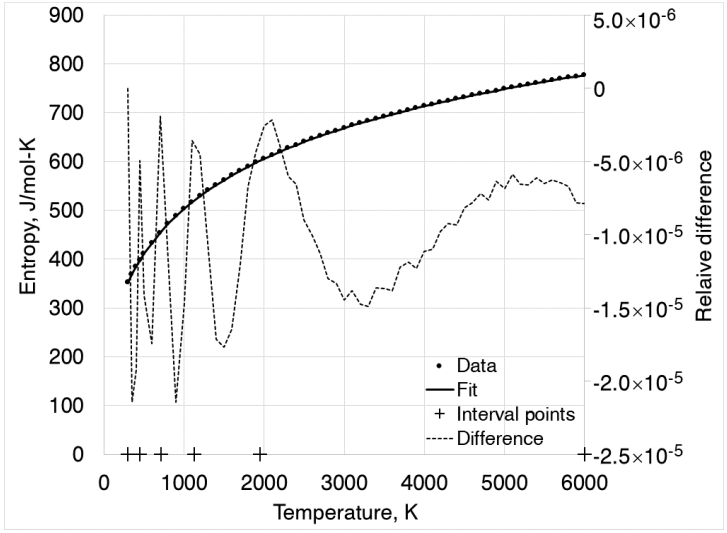


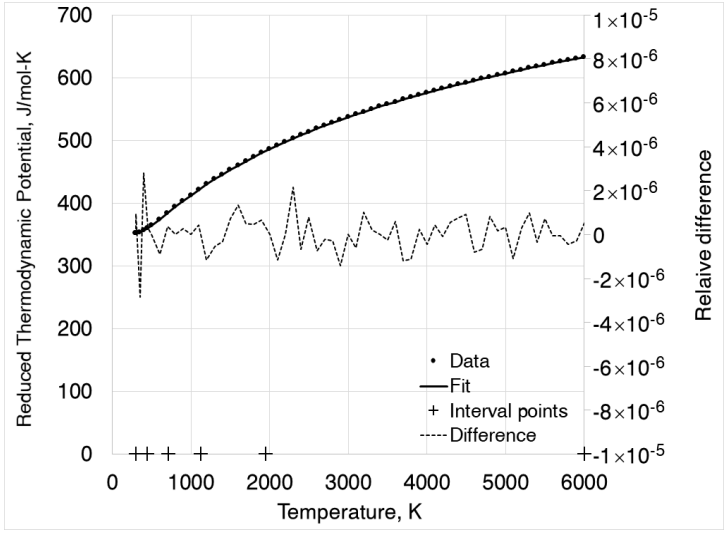
Tungstic Acid ($O_2W(OH)_2$) – H2O4W1(g)

MW = 249.86 g/mol – $\Delta_f H^o$(298.15 K) = -905.836 kJ/mol – S^o(298.15 K) = 351.999 J/mol-K – p^o = 0.1 MPa



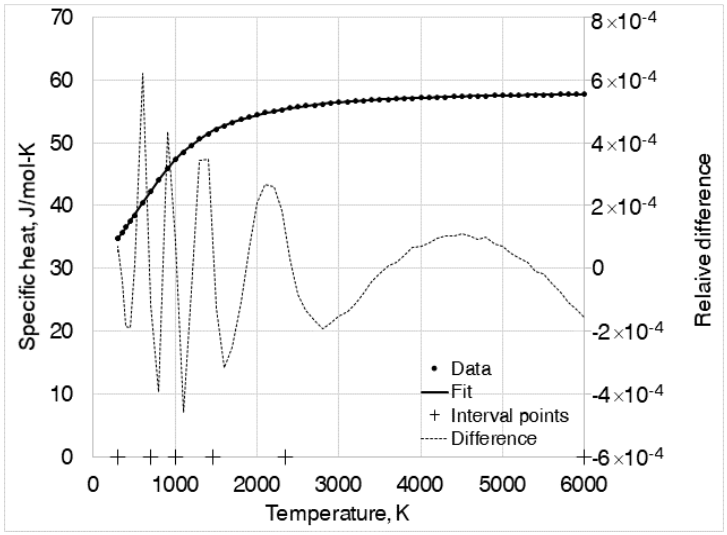


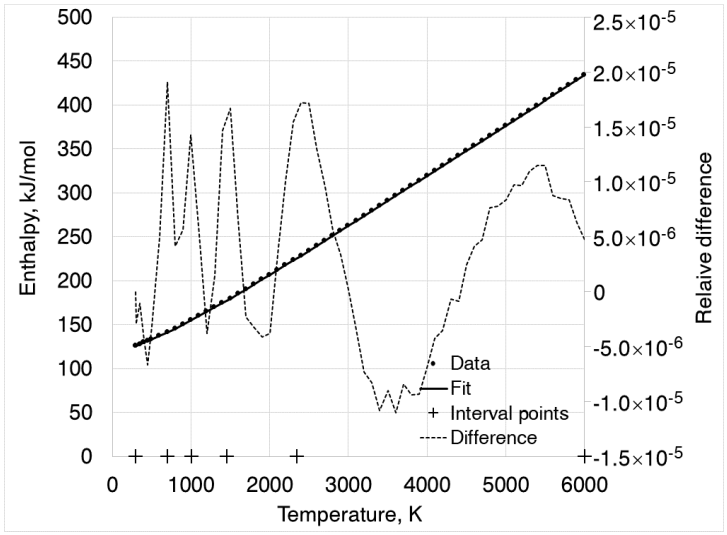


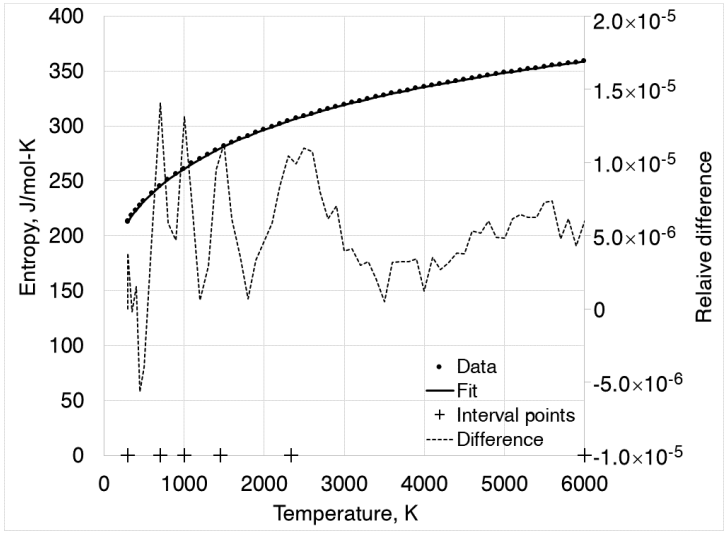


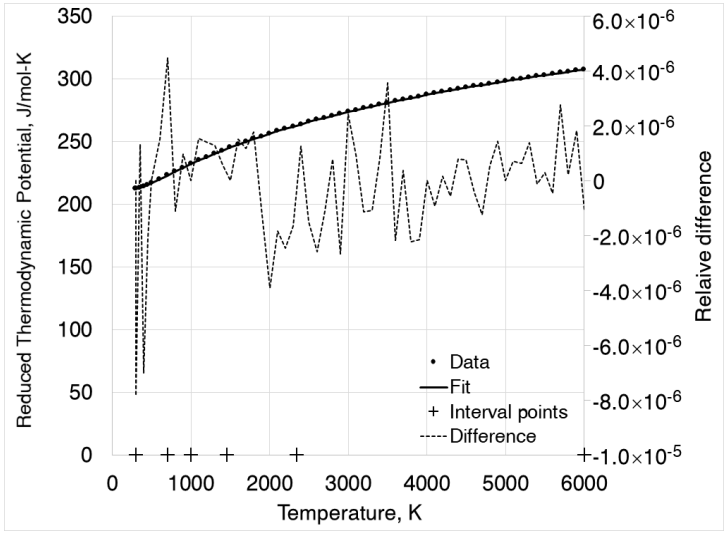
Phosphino (PH_2) – H2P1(g)

MW = 32.98964 g/mol – $\Delta_f H^o$(298.15 K) = 125.938 kJ/mol – S^o(298.15 K) = 212.652 J/mol-K – p^o = 0.1 MPa



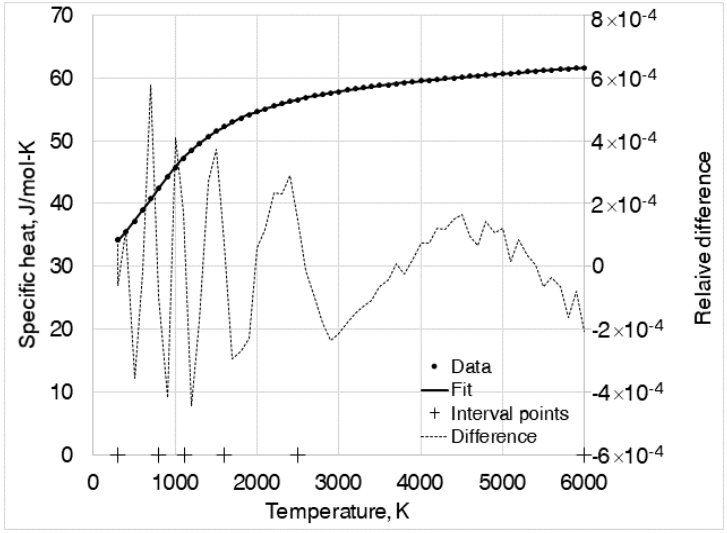


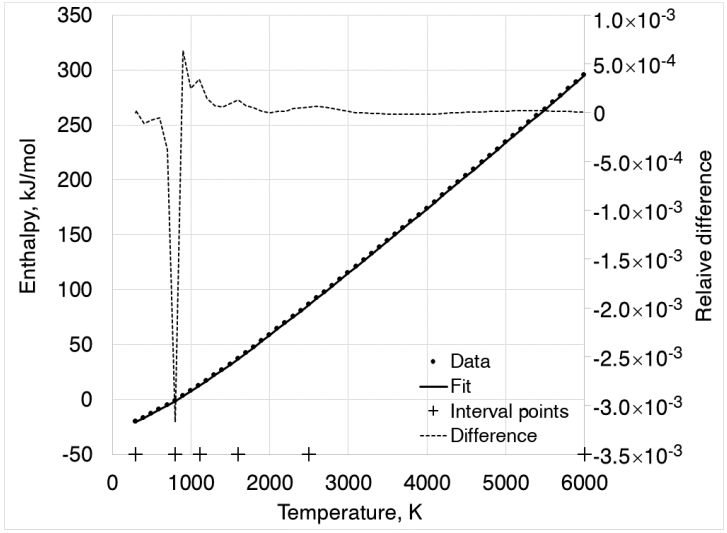


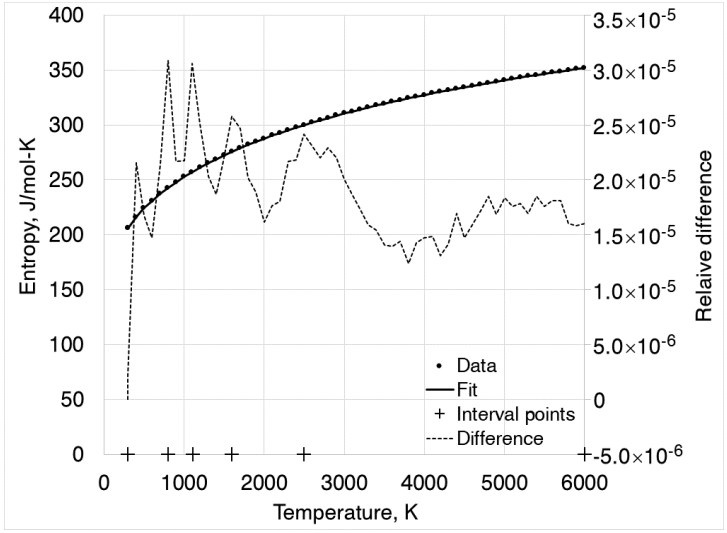


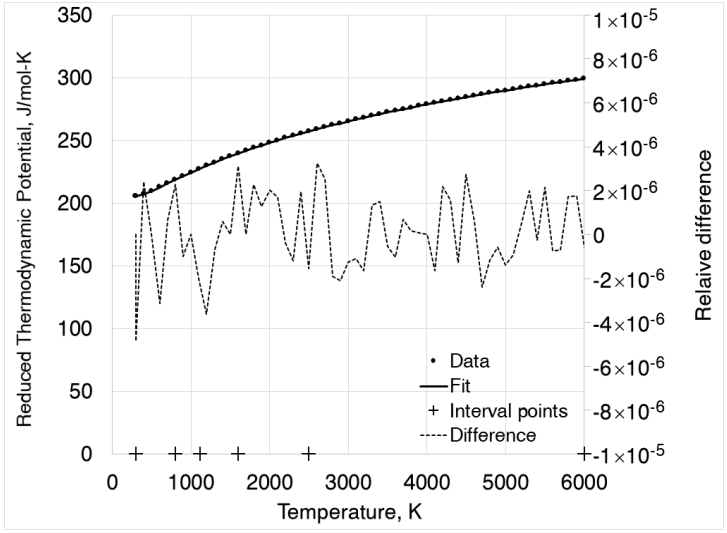
Hydrogen Sulfide (H_2S) – H2S1(g)

MW = 34.08 g/mol – $\Delta_f H^o$(298.15 K) = -20.502 kJ/mol – S^o(298.15 K) = 205.757 J/mol-K – p^o = 0.1 MPa



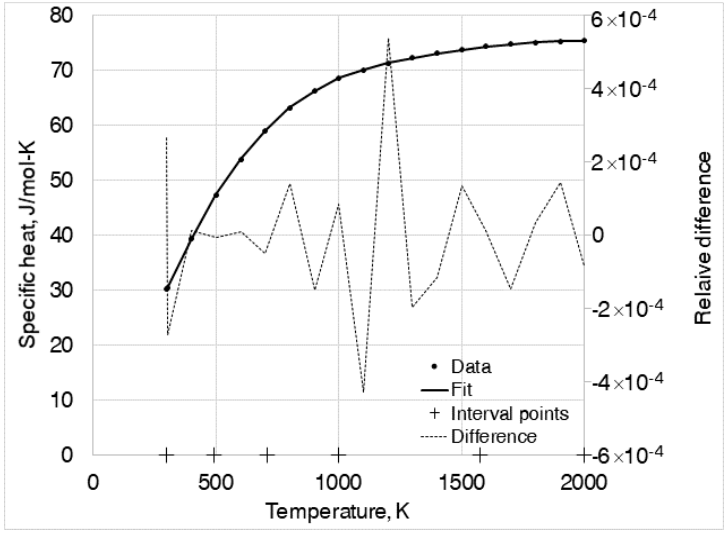


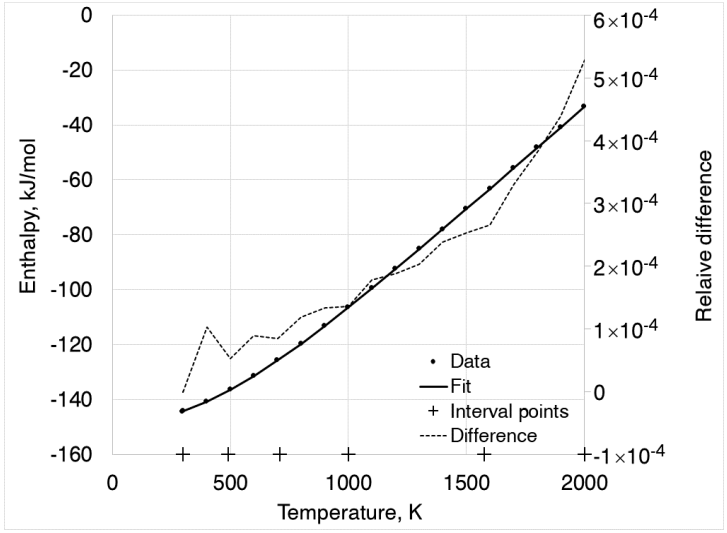


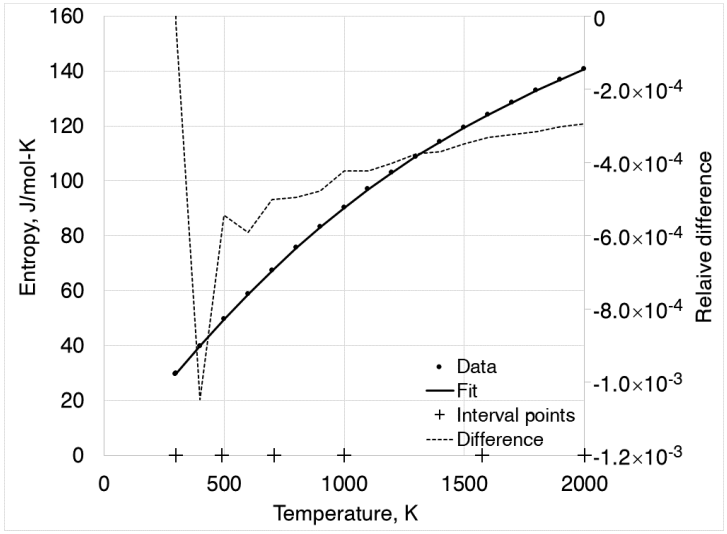


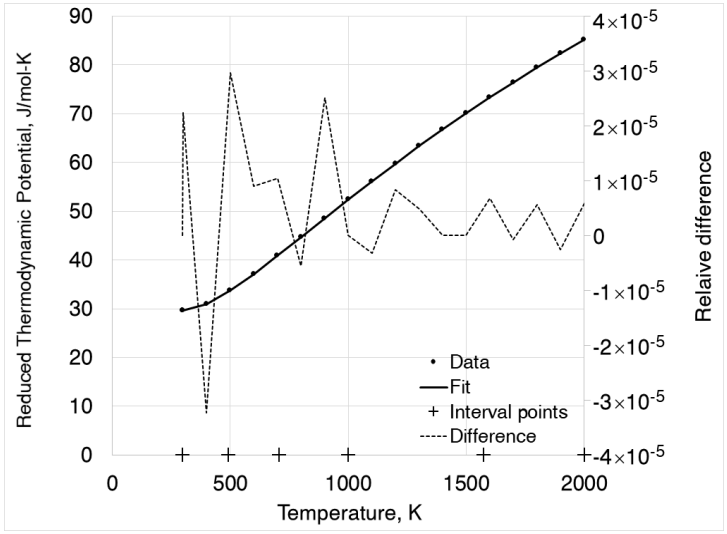
Titanium Hydride (TiH_2) – H2Ti1(cr)

MW = 49.9 g/mol – $\Delta_f H^o$(298.15 K) = -144.348 kJ/mol – S^o(298.15 K) = 29.711 J/mol-K – p^o = 0.1 MPa



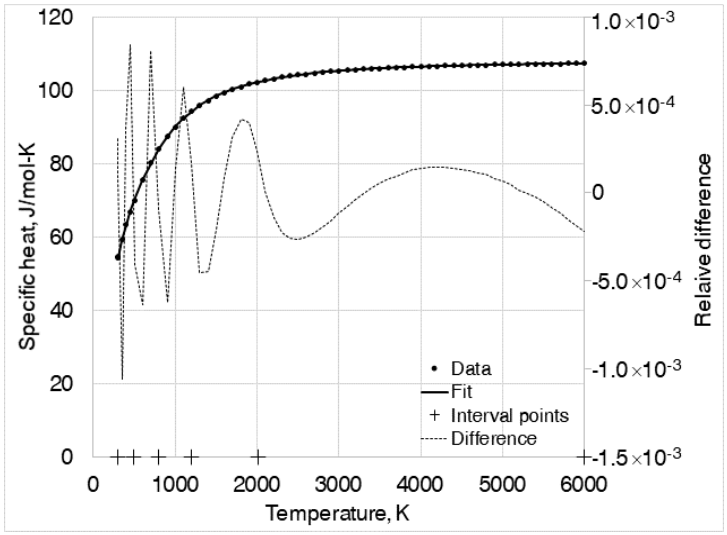


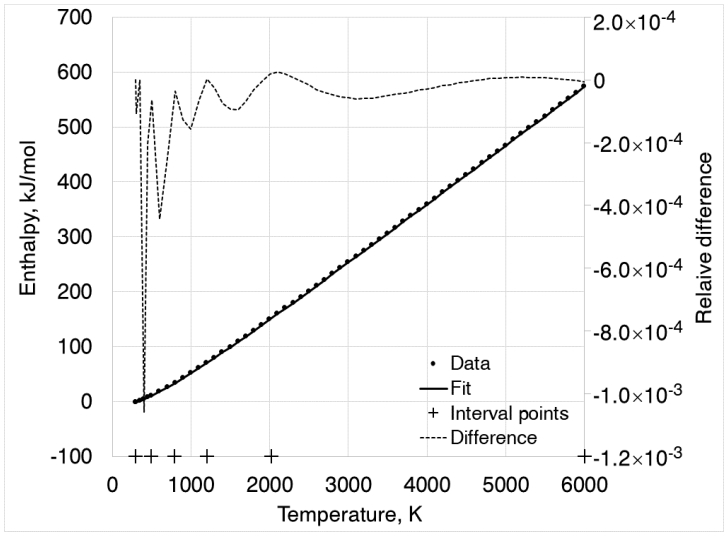


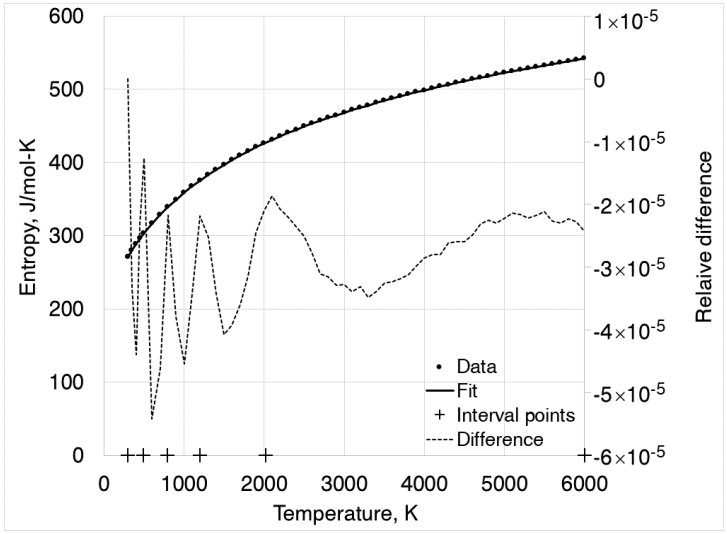


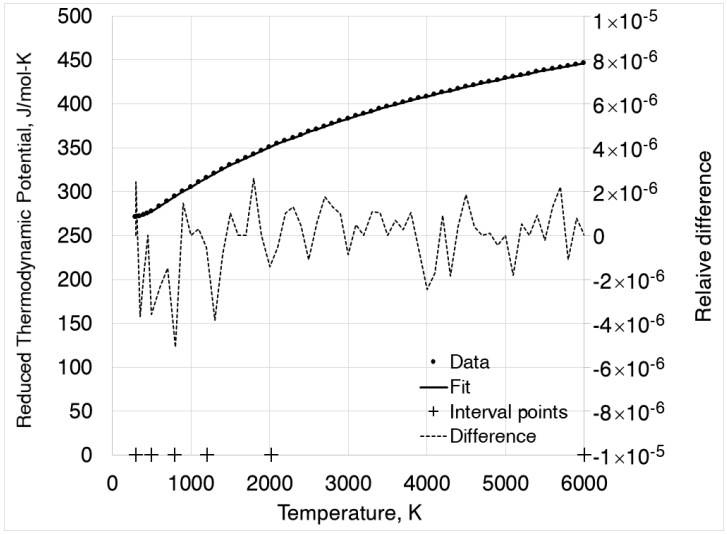
Iodosilane (SiH_3I) – H3I1Si1(g)

MW = 158.0138 g/mol – $\Delta_f H^o$(298.15 K) = -2.092 kJ/mol – S^o(298.15 K) = 271.022 J/mol-K – p^o = 0.1 MPa



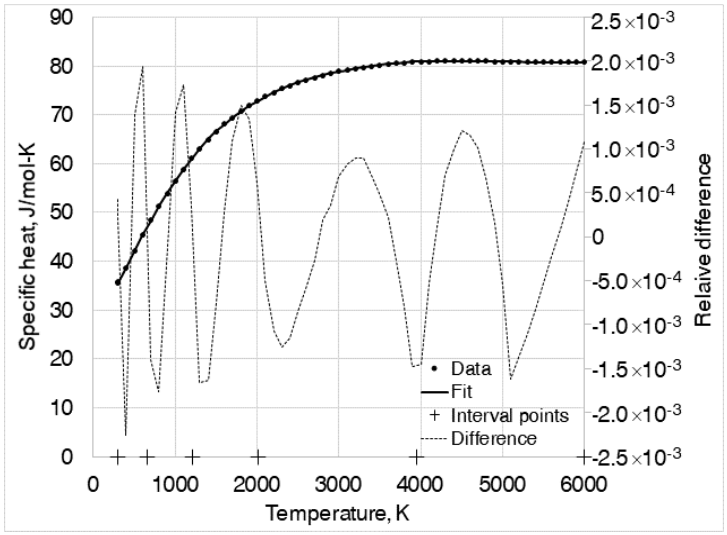


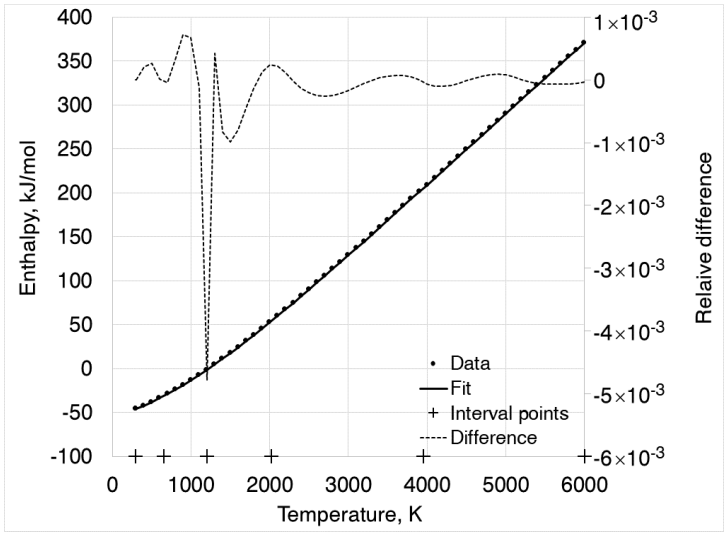


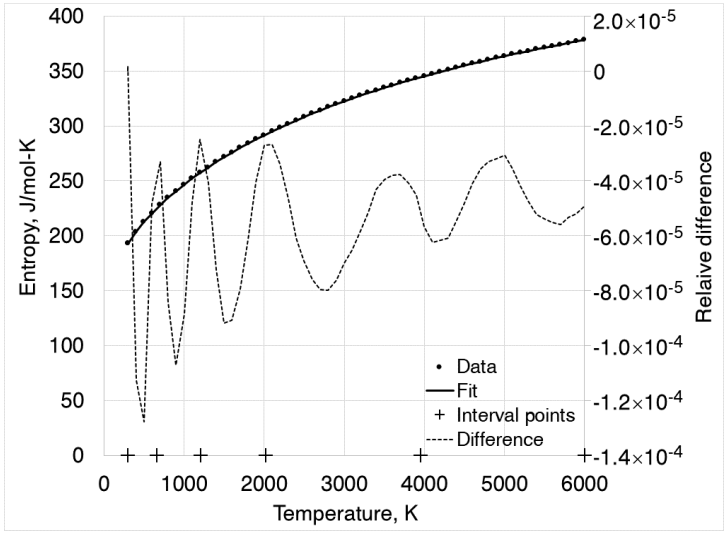


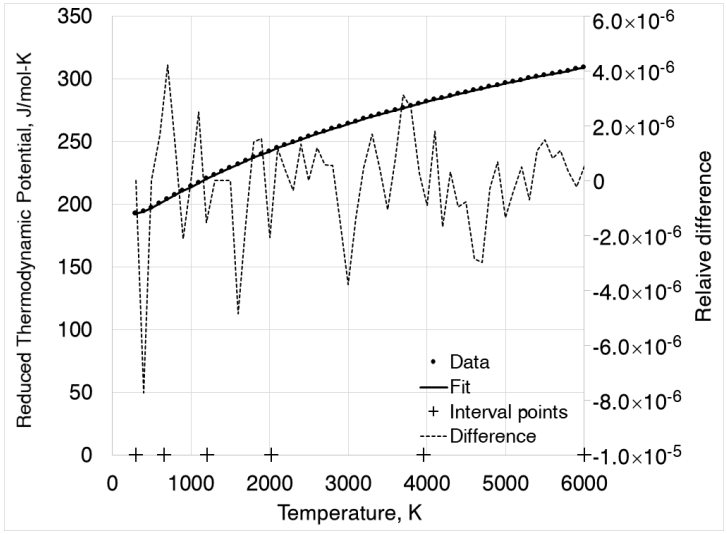
Ammonia (NH_3) – H3N1(g)

MW = 17.0305 g/mol – $\Delta_f H^o$(298.15 K) = -45.898 kJ/mol – S^o(298.15 K) = 192.774 J/mol-K – p^o = 0.1 MPa



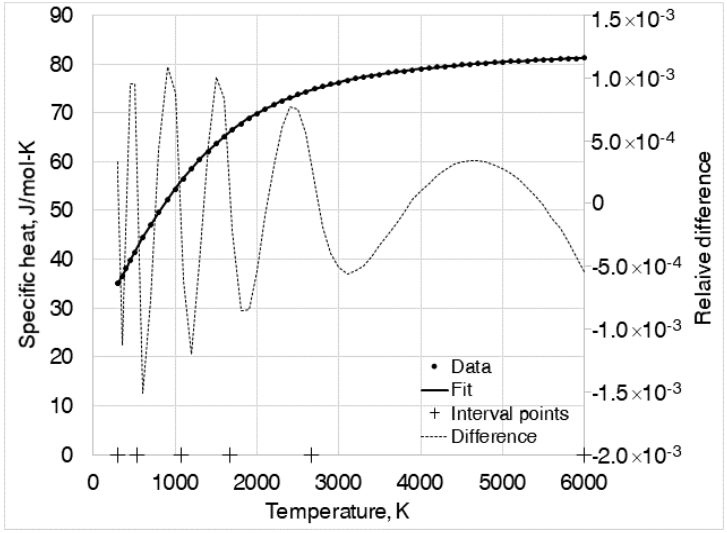


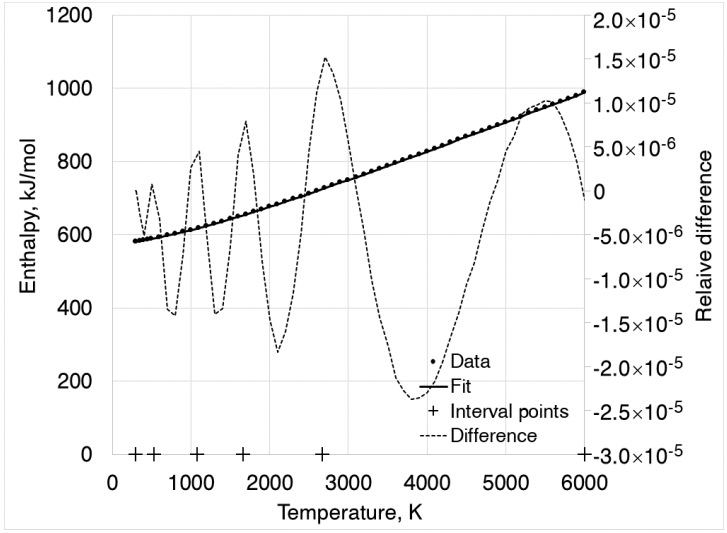


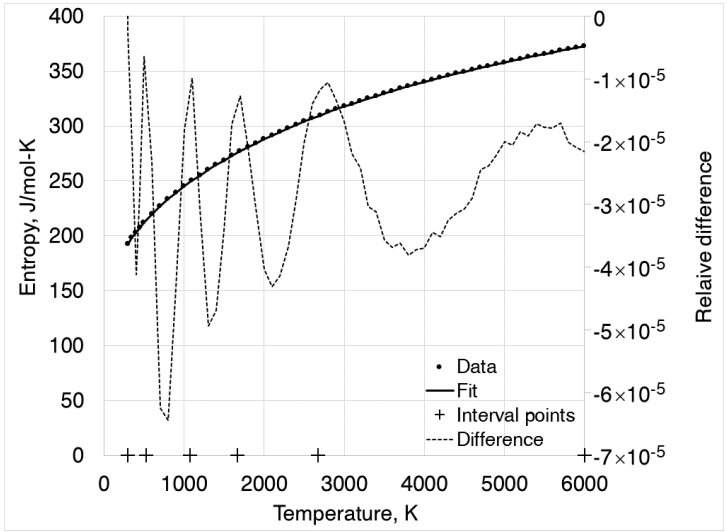


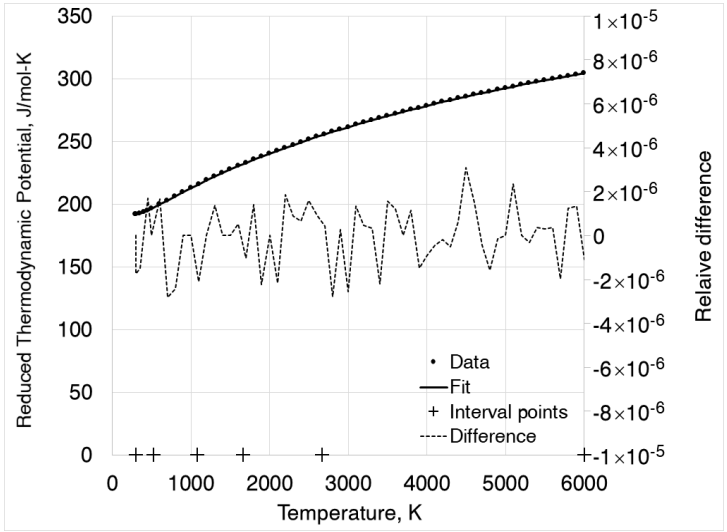
Hydronium, Ion (H_3O^+) – H3O1+(g)

MW = 19.0227 g/mol – $\Delta_f H^o$(298.15 K) = 581.158 kJ/mol – S^o(298.15 K) = 192.257 J/mol-K – p^o = 0.1 MPa



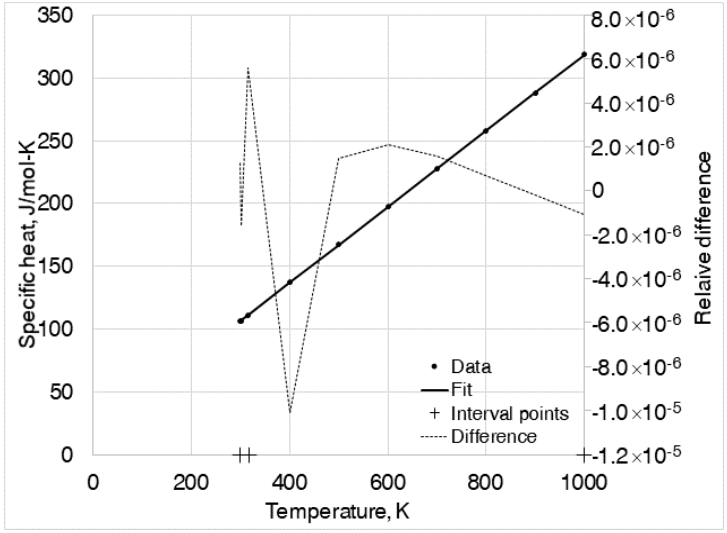


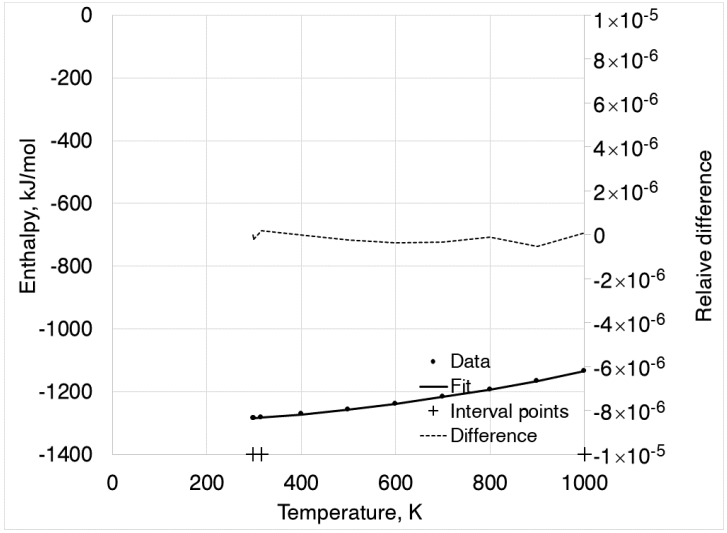


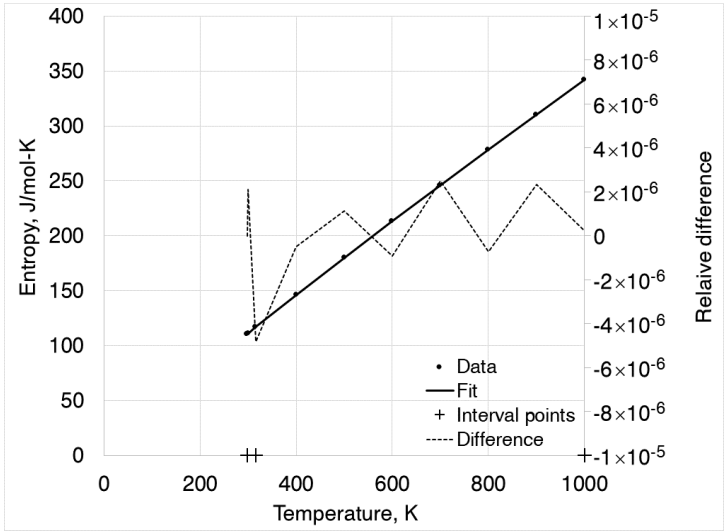


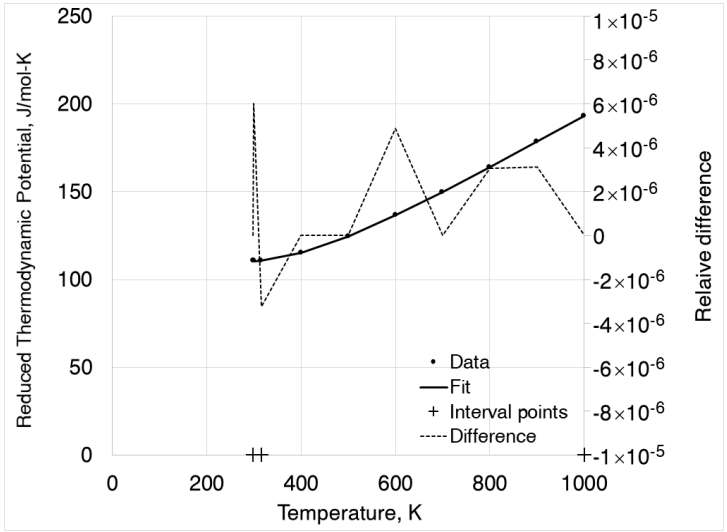
Phosphoric Acid (H_3PO_4) – H3O4P1(cr)

MW = 97.9952 g/mol – $\Delta_f H^o$(298.15 K) = -1284.375 kJ/mol – S^o(298.15 K) = 110.544 J/mol-K – p^o = 0.1 MPa



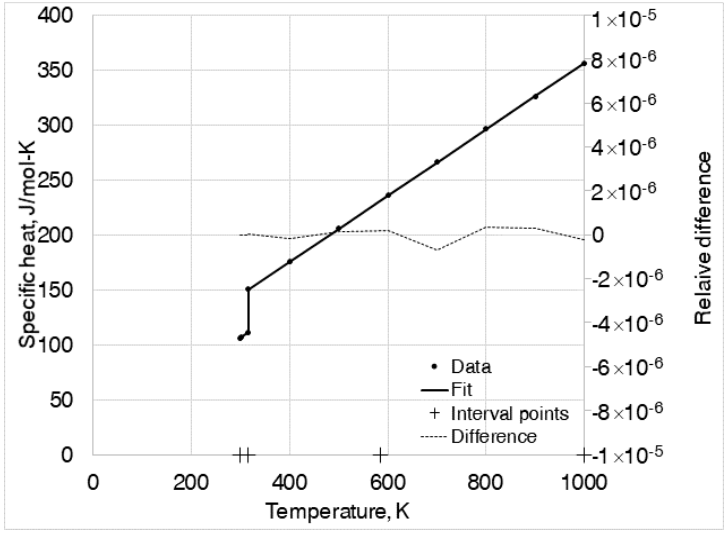


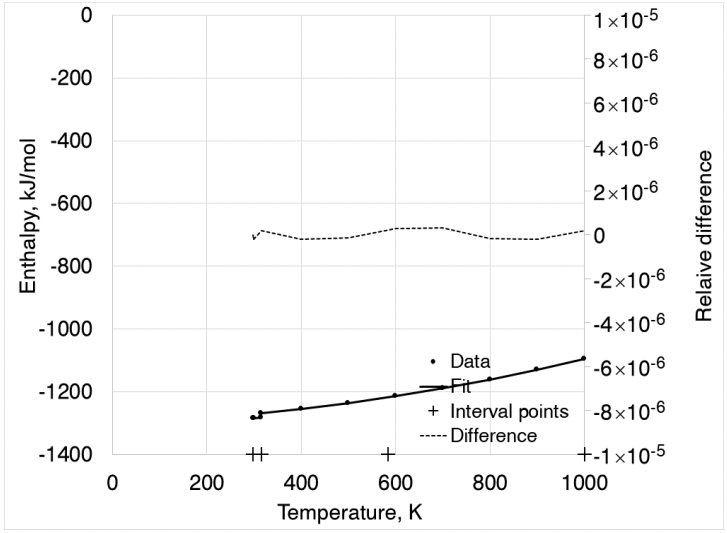


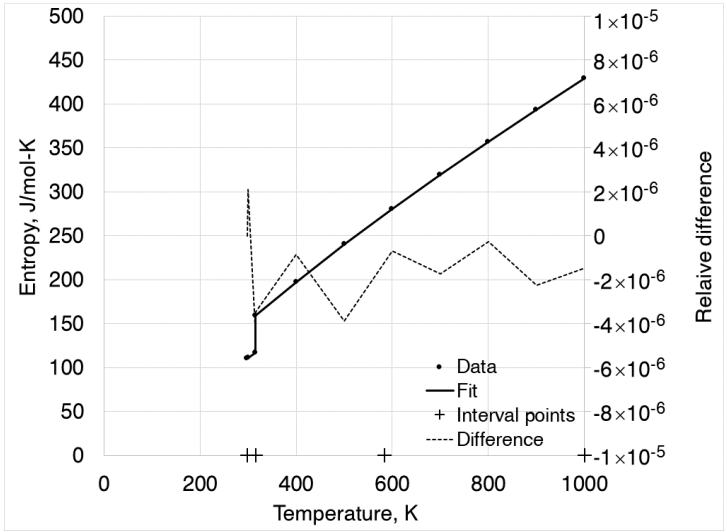


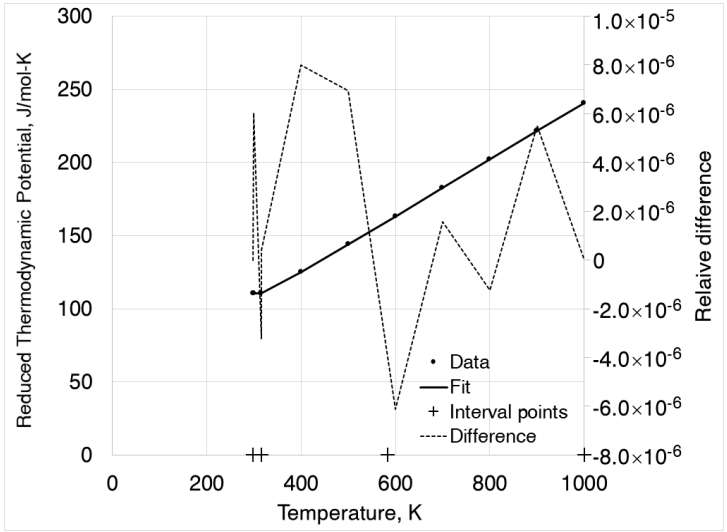
Phosphoric Acid (H_3PO_4) – H3O4P1(cr,l)

MW = 97.9952 g/mol – $\Delta_f H^o$(298.15 K) = -1284.375 kJ/mol – S^o(298.15 K) = 110.544 J/mol-K – p^o = 0.1 MPa



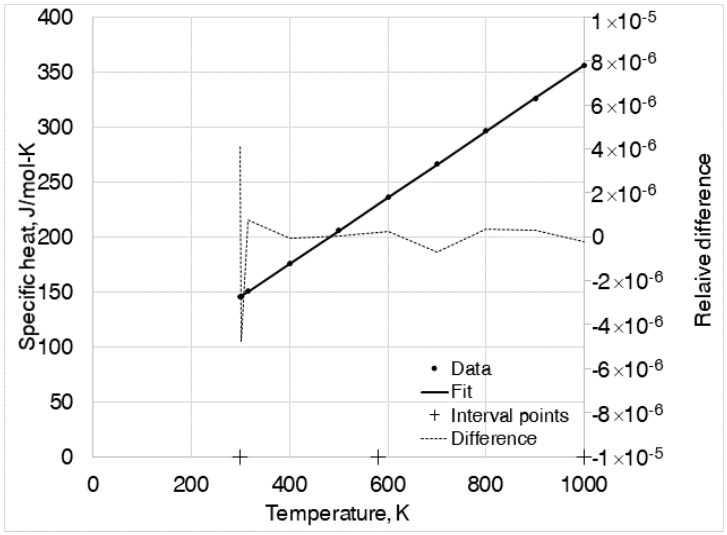


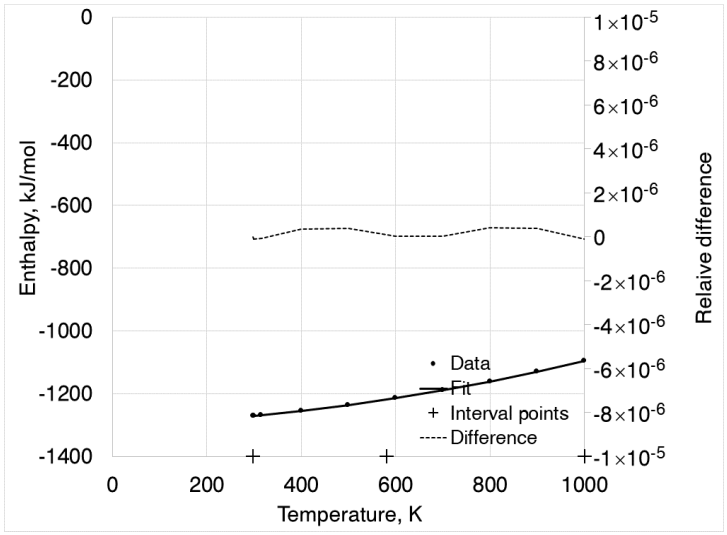


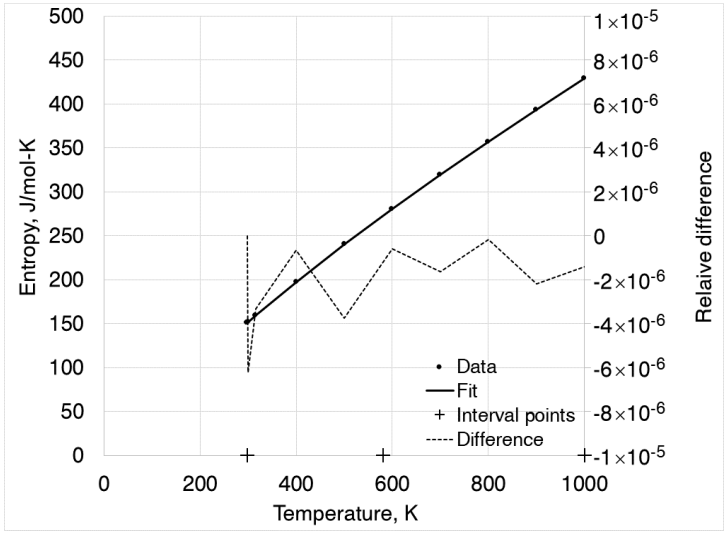


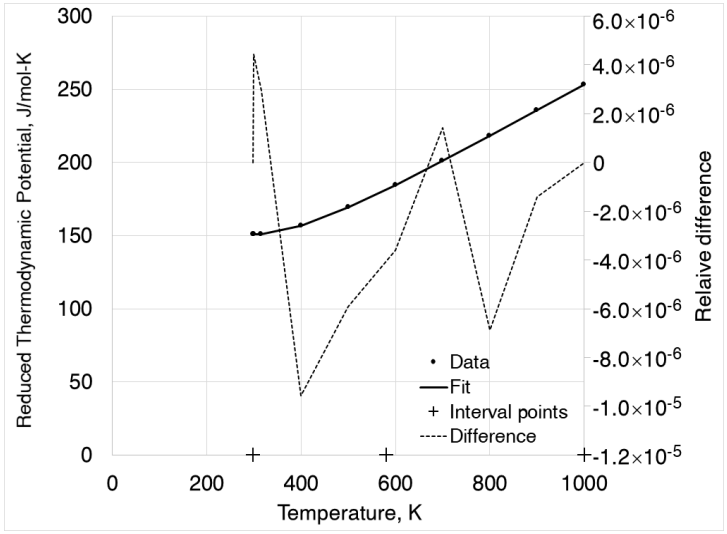
Phosphoric Acid (H_3PO_4) – H3O4P1(l)

MW = 97.9952 g/mol – $\Delta_f H^o$(298.15 K) = -1271.663 kJ/mol – S^o(298.15 K) = 150.775 J/mol-K – p^o = 0.1 MPa



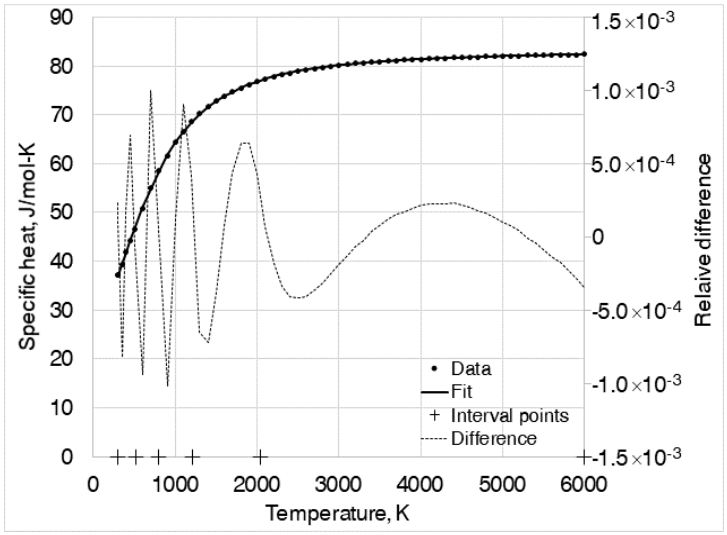


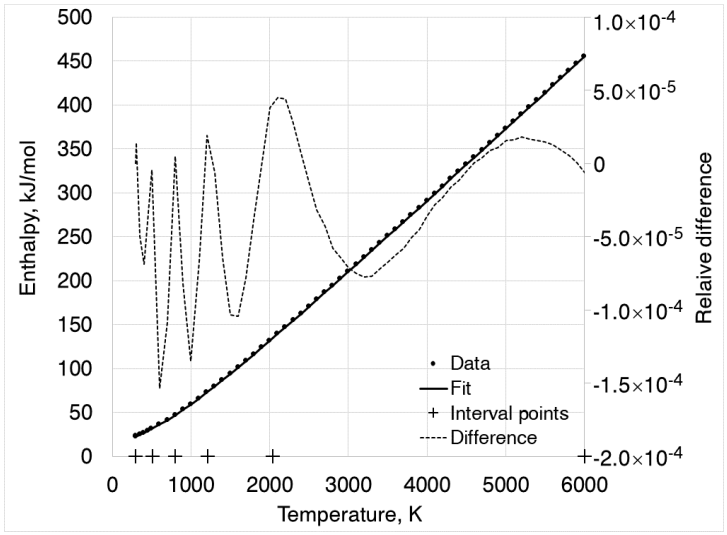


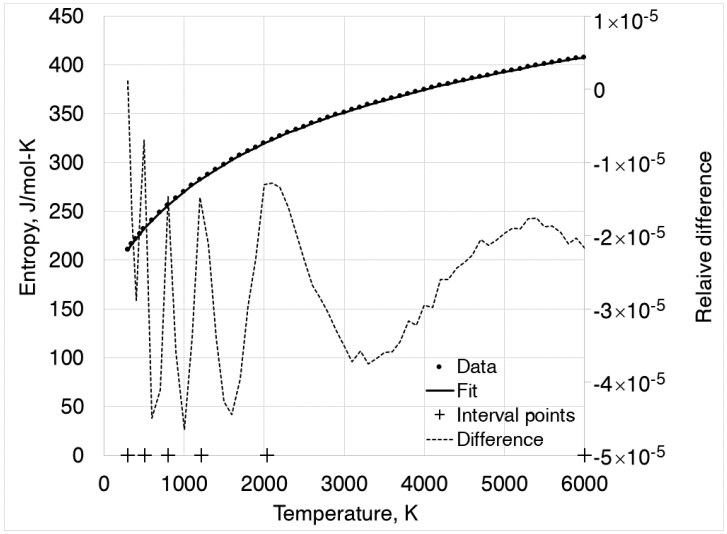


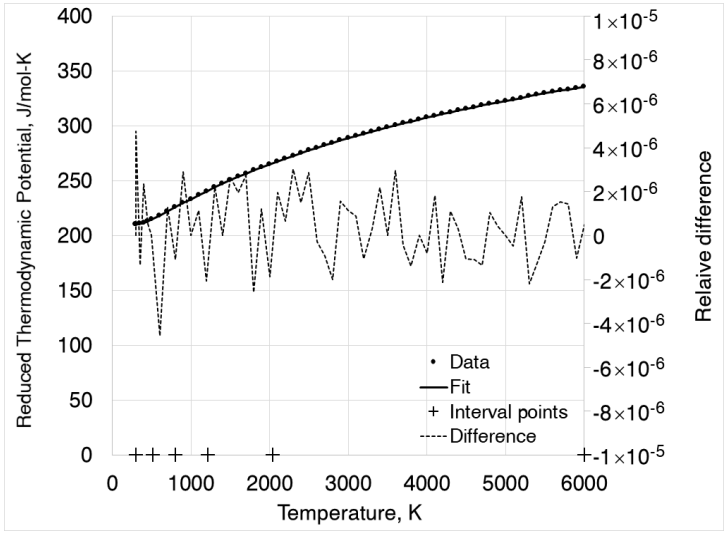
Phosphine (PH_3) – H3P1(g)

MW = 33.99758 g/mol – $\Delta_f H^o$(298.15 K) = 22.886 kJ/mol – S^o(298.15 K) = 210.243 J/mol-K – p^o = 0.1 MPa



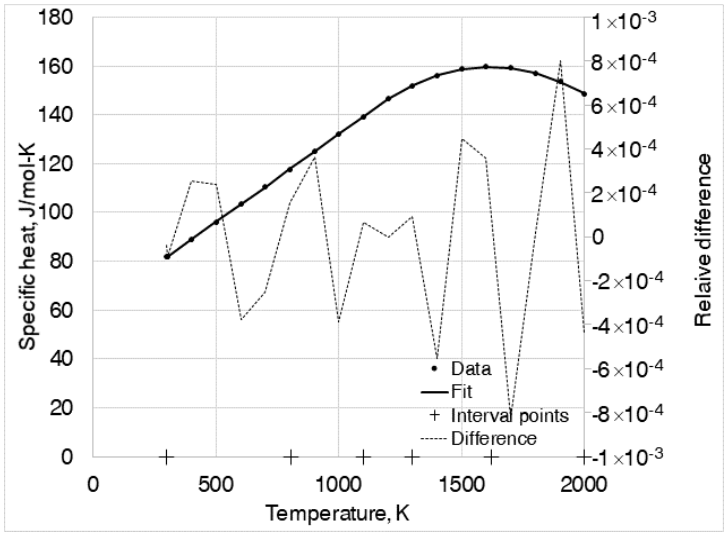


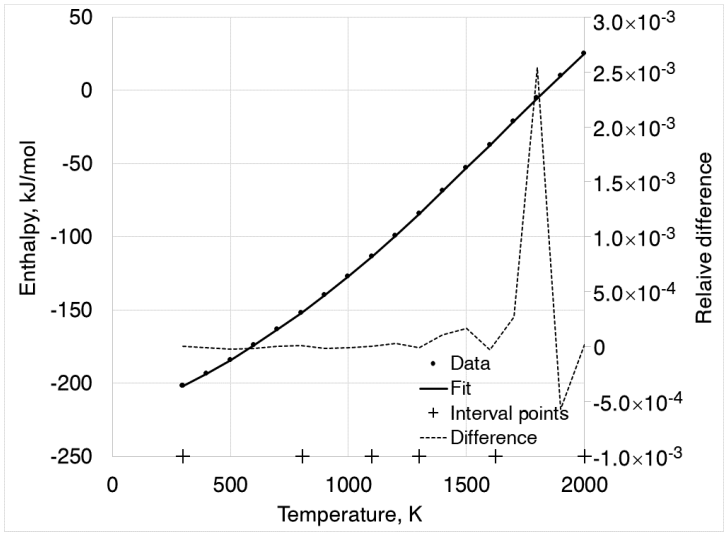


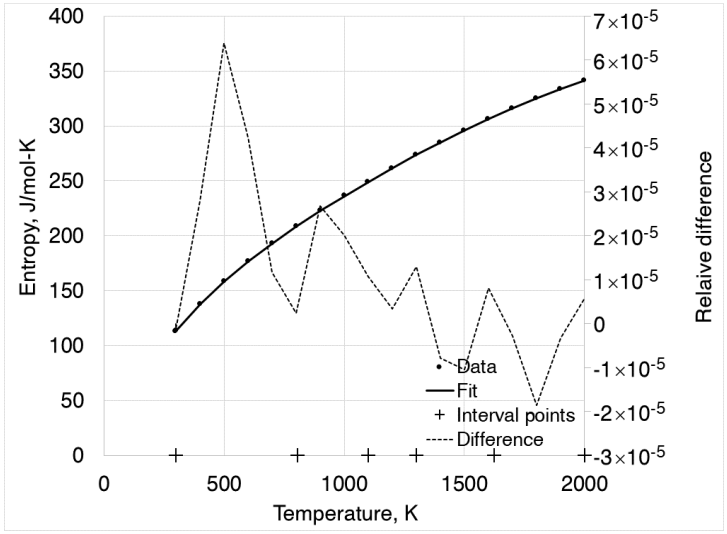


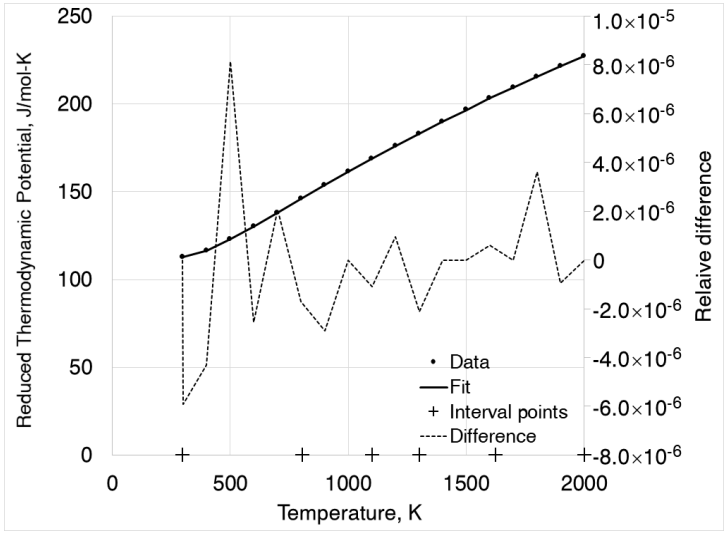
Ammonium Iodide (NH_4I) – H4I1N1(cr)

MW = 144.943 g/mol – $\Delta_f H^o$(298.15 K) = -202.087 kJ/mol – S^o(298.15 K) = 112.968 J/mol-K – p^o = 0.1 MPa



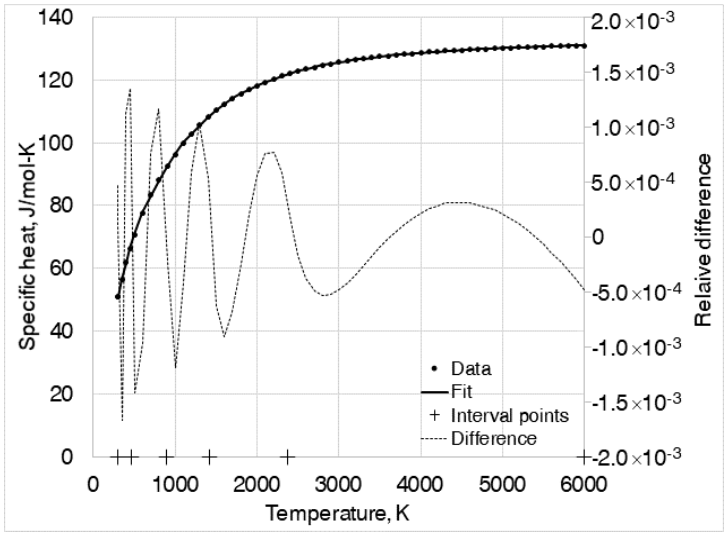


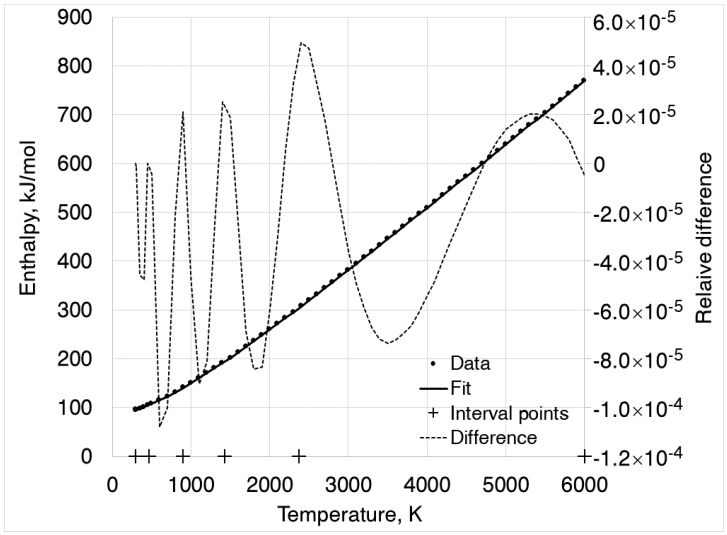


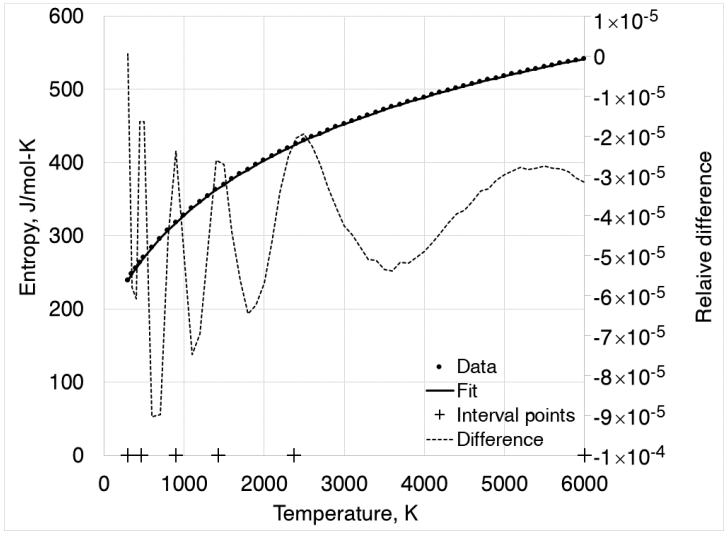


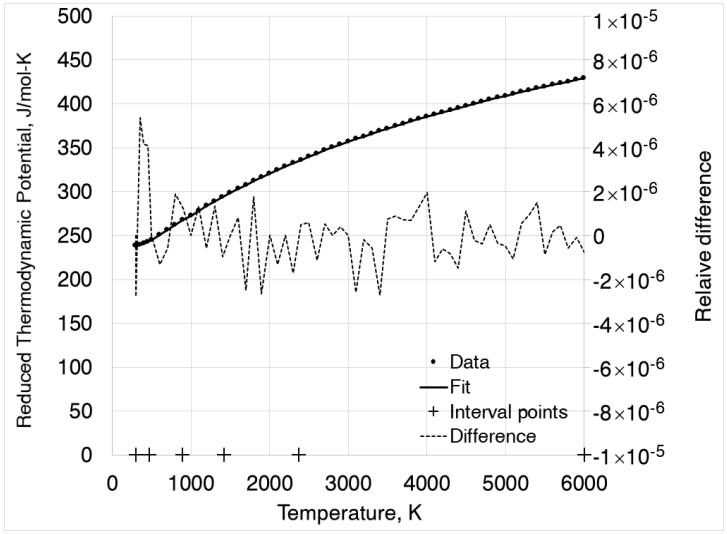
Hydrazine (N_2H_4) – H4N2(g)

MW = 32.0452 g/mol – $\Delta_f H^o$(298.15 K) = 95.353 kJ/mol – S^o(298.15 K) = 238.719 J/mol-K – p^o = 0.1 MPa



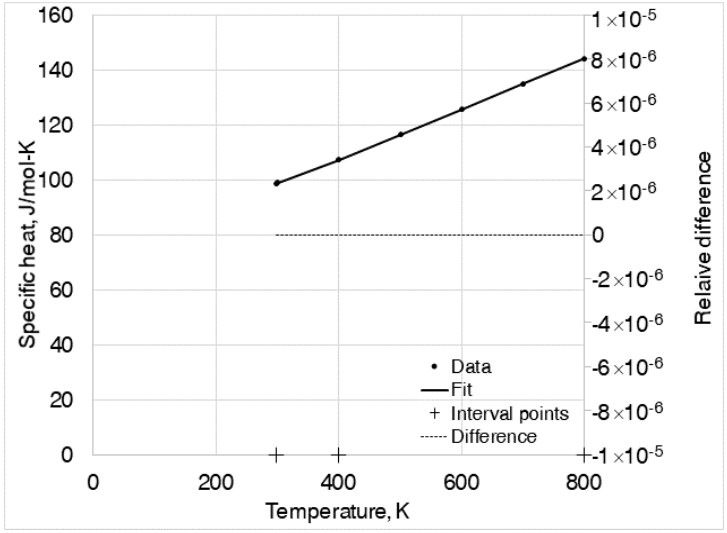


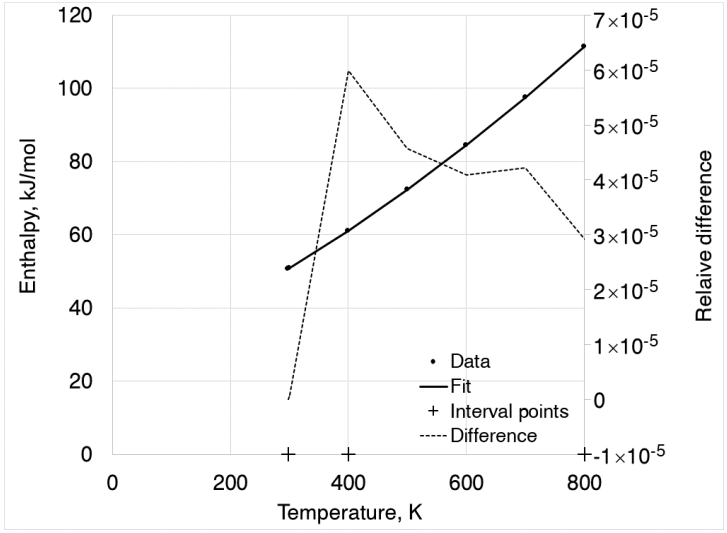


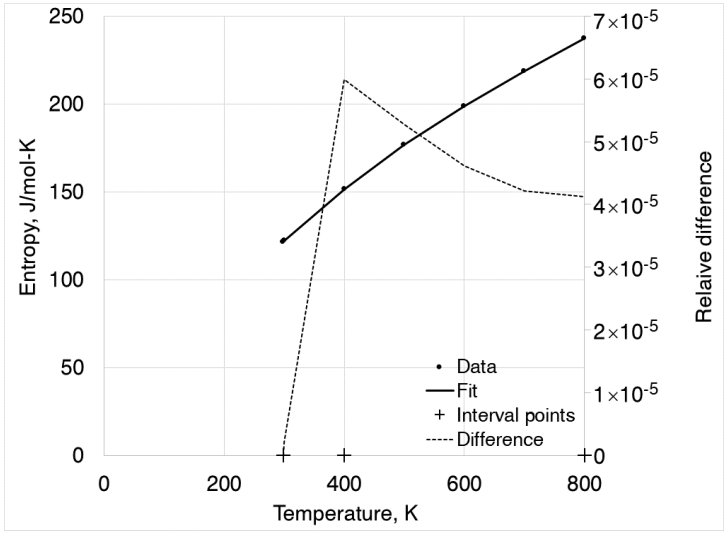


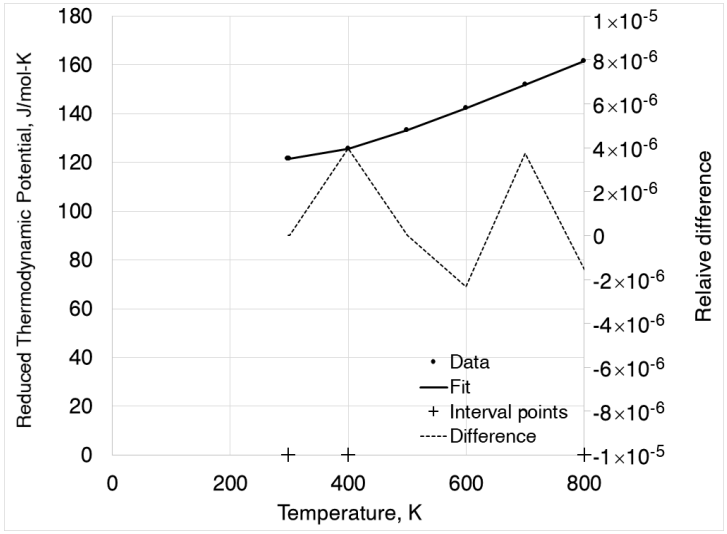
Hydrazine (N_2H_4) – H4N2(l)

MW = 32.0452 g/mol – $\Delta_f H^o$(298.15 K) = 50.626 kJ/mol – S^o(298.15 K) = 121.544 J/mol-K – p^o = 0.1 MPa



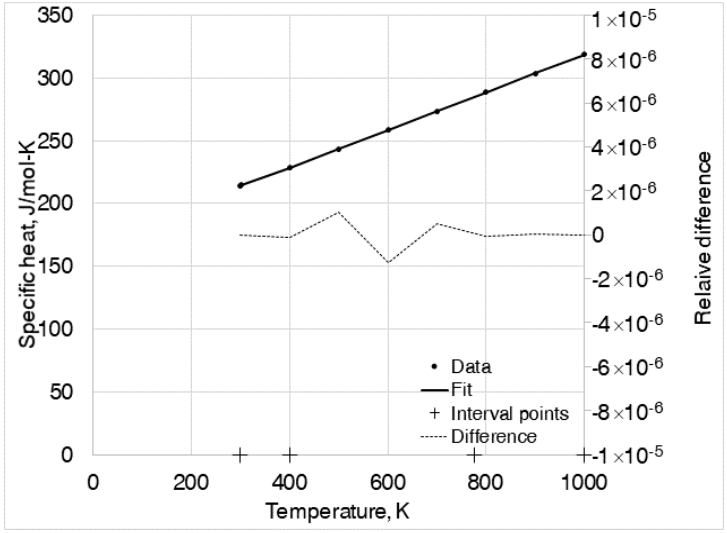


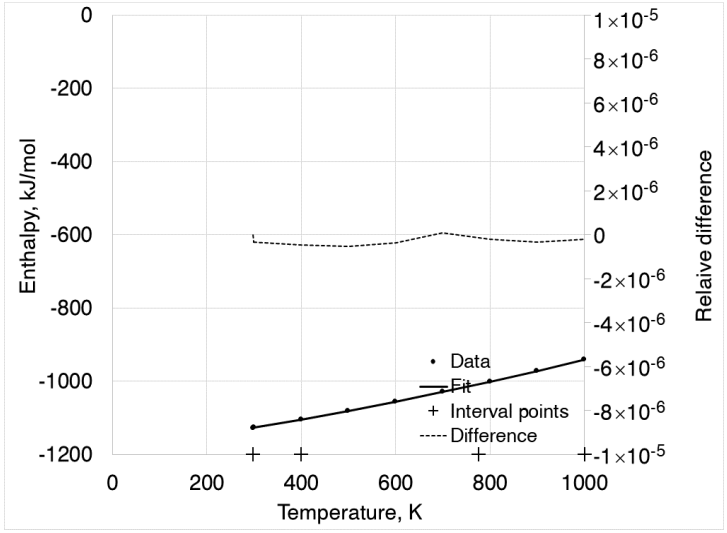


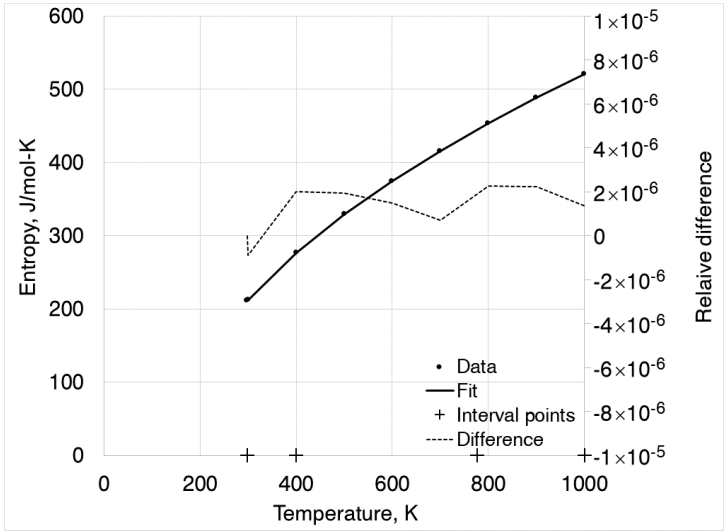


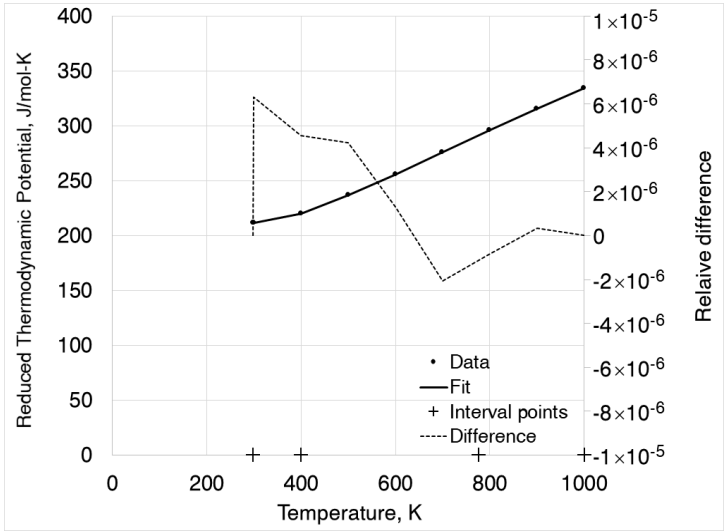
Sulfuric Acid, Monohydrate ($H_2SO_4.H_2O$) – H4O5S1(cr,l)

MW = 116.09 g/mol – $\Delta_f H^o$(298.15 K) = -1127.621 kJ/mol – S^o(298.15 K) = 211.51 J/mol-K – p^o = 0.1 MPa



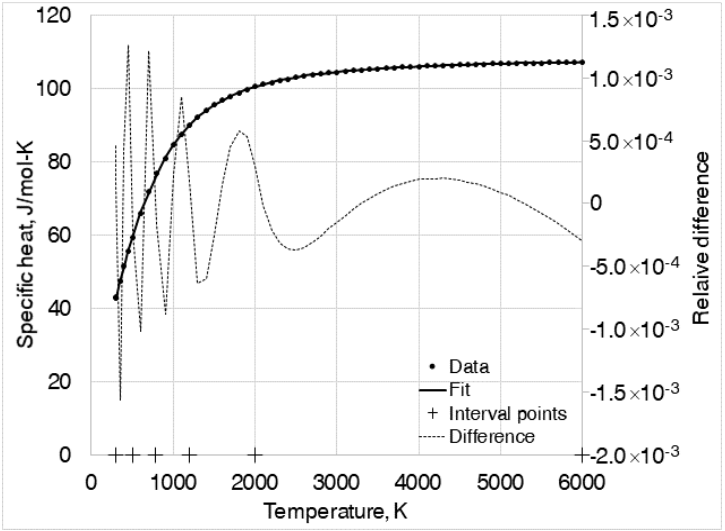


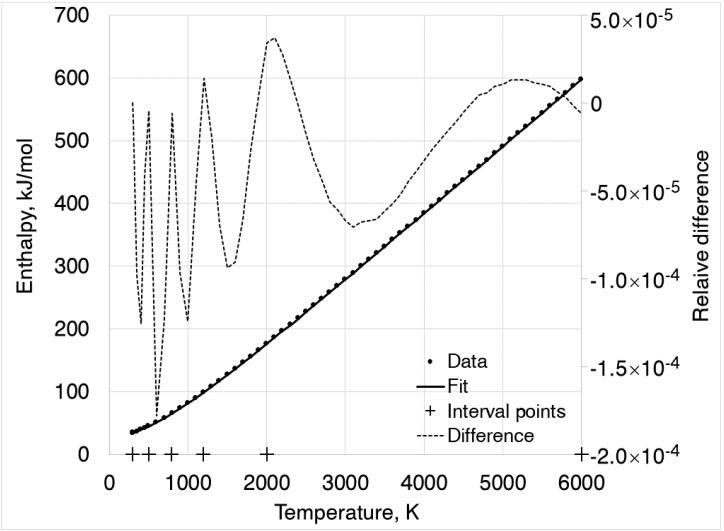


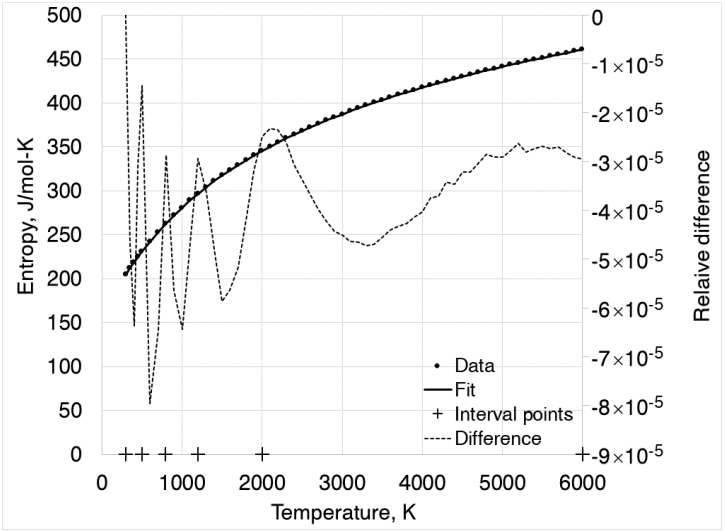


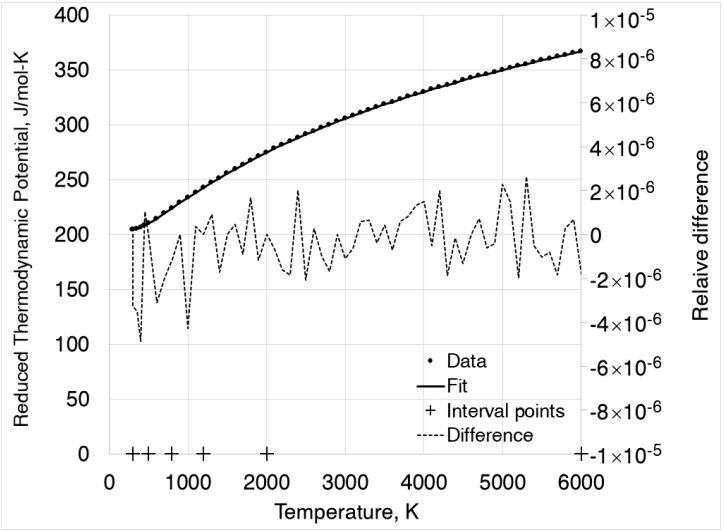
Silane (SiH_4) – H4Si1(g)

MW = 32.1173 g/mol – $\Delta_f H^o$(298.15 K) = 34.309 kJ/mol – S^o(298.15 K) = 204.653 J/mol-K – p^o = 0.1 MPa



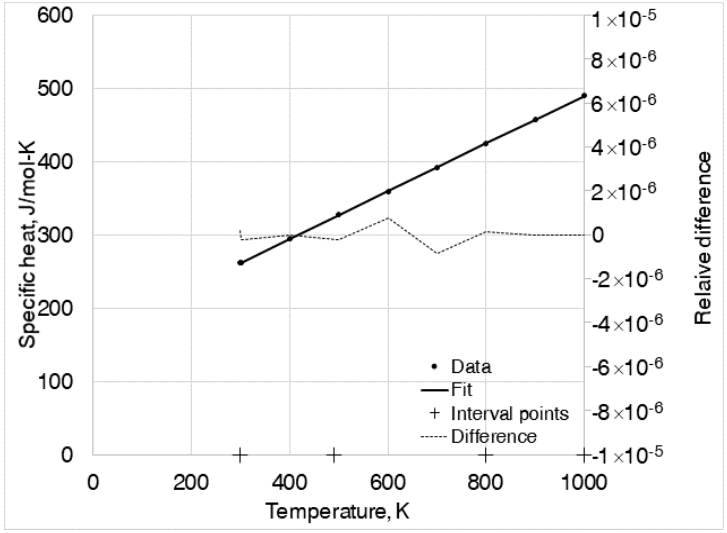


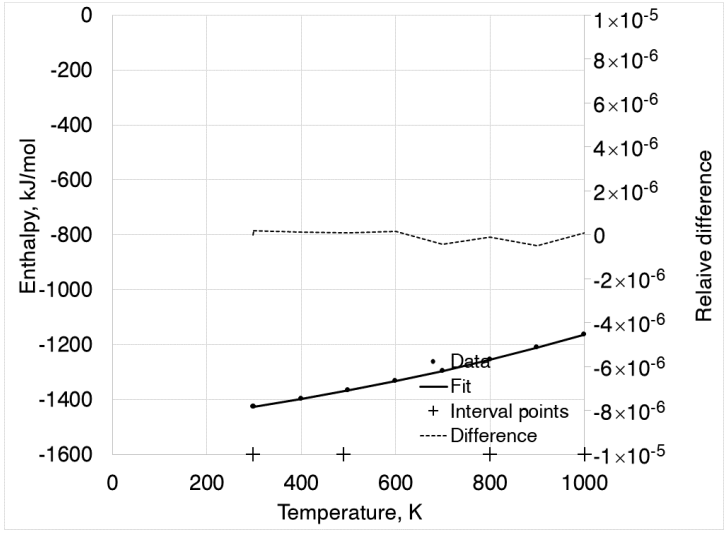


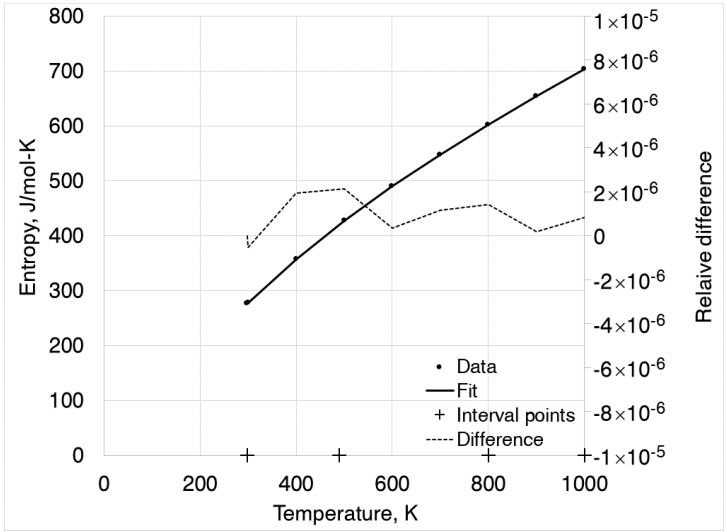


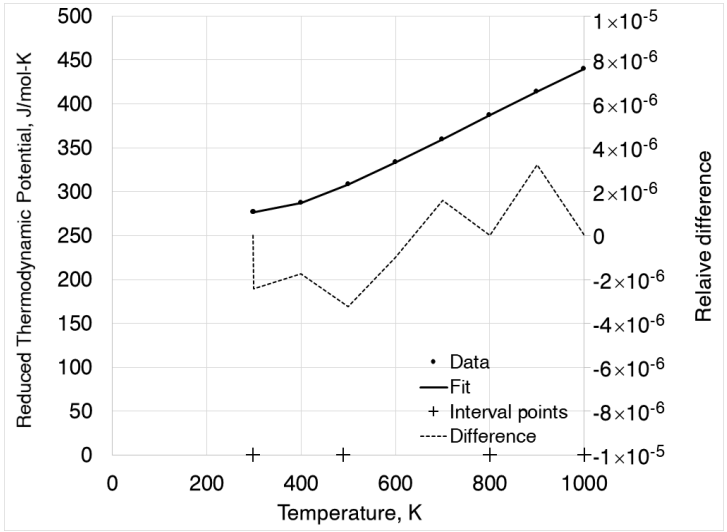
Sulfuric Acid, Dihydrate ($H_2SO_4.2H_2O$) – H6O6S1(cr,l)

MW = 134.1 g/mol – $\Delta_f H^o$(298.15 K) = -1427.1 kJ/mol – S^o(298.15 K) = 276.363 J/mol-K – p^o = 0.1 MPa



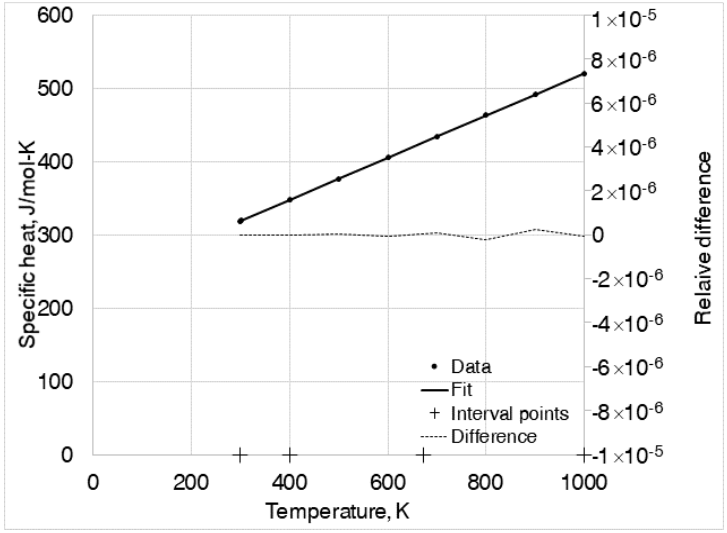


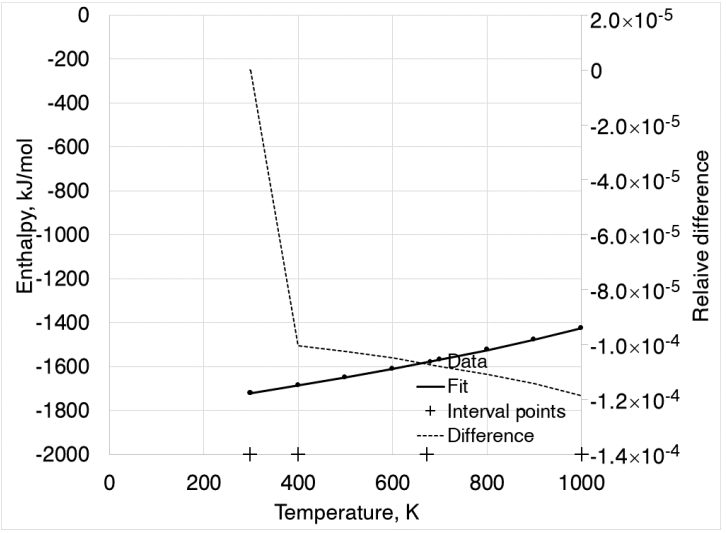


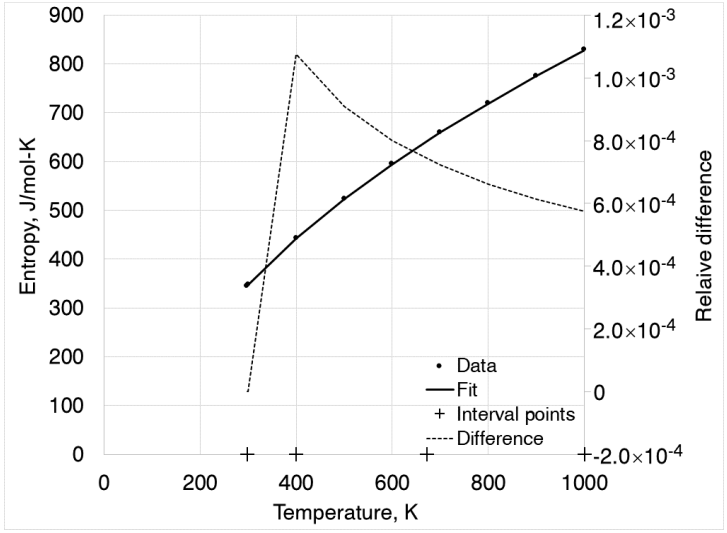


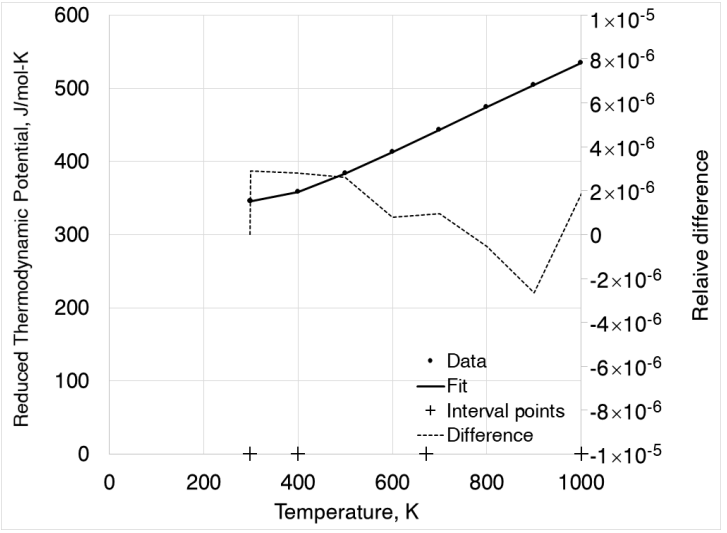
Sulfuric Acid, Trihydrate ($H_2SO_4.3H_2O$) – H8O7S1(cr,l)

MW = 152.12 g/mol – $\Delta_f H^o$(298.15 K) = -1720.402 kJ/mol – S^o(298.15 K) = 345.373 J/mol-K – p^o = 0.1 MPa



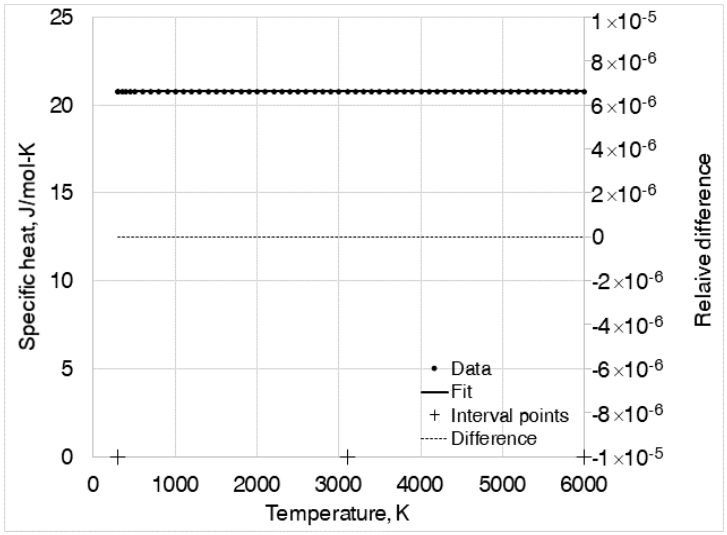


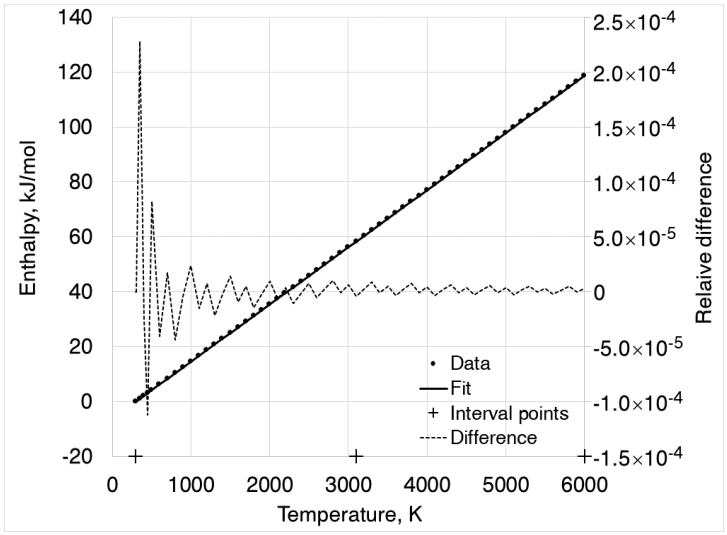


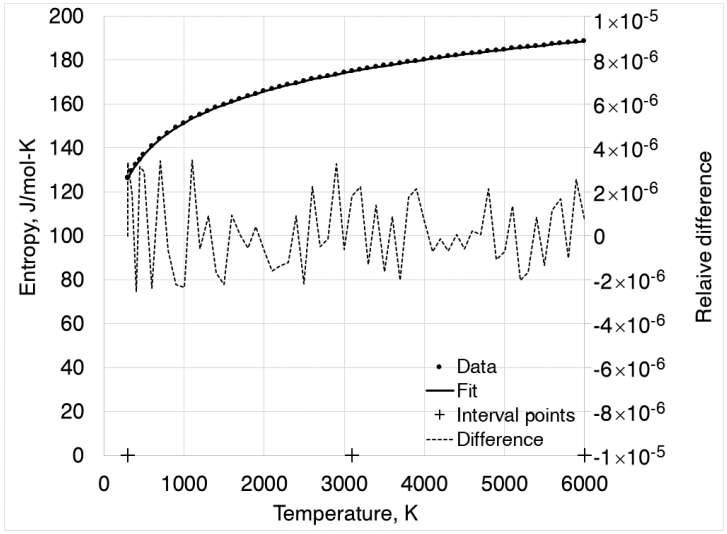


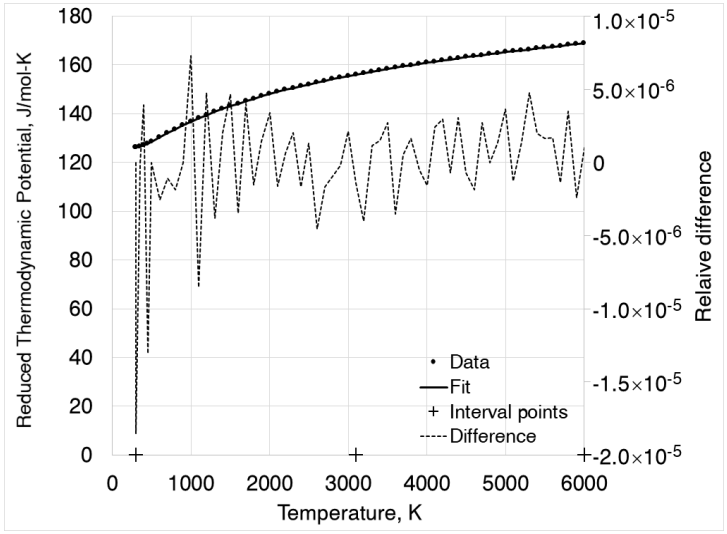
Helium (He) – He1(ref)

MW = 4.0026 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 126.152 J/mol-K – p^o = 0.1 MPa



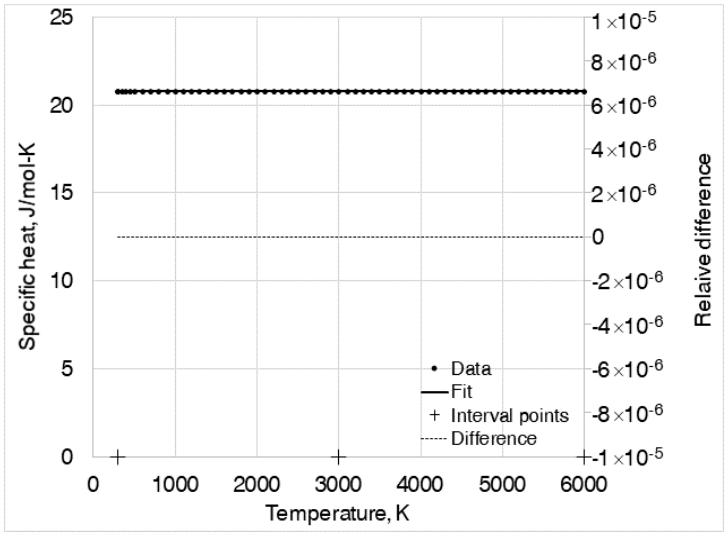


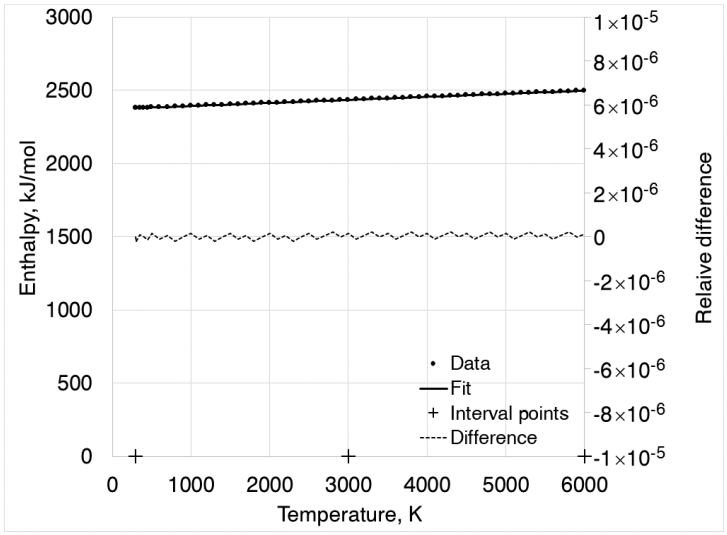


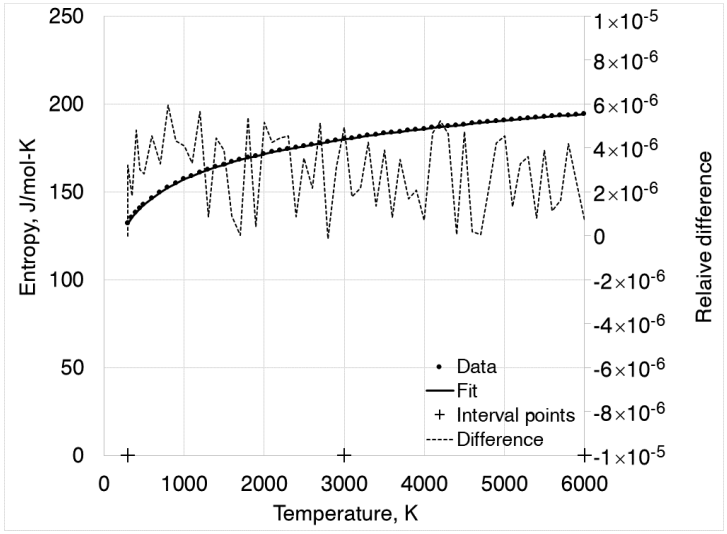


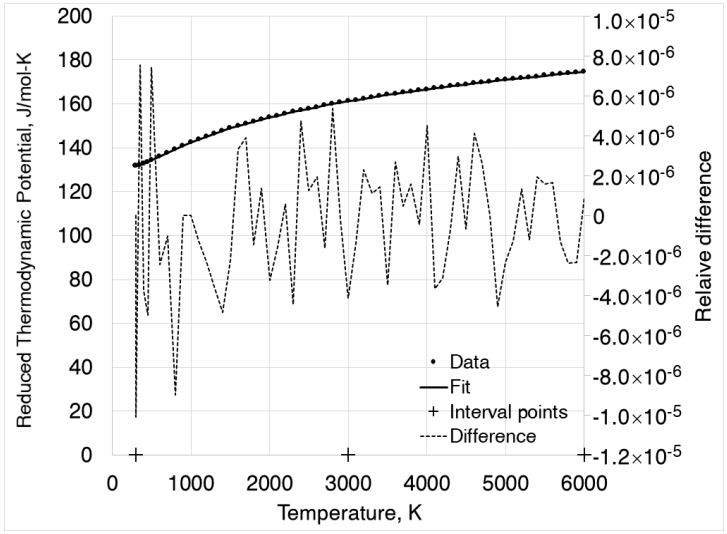
Helium, Ion (He^+) – He1+(g)

MW = 4.00205 g/mol – $\Delta_f H^o$(298.15 K) = 2378.522 kJ/mol – S^o(298.15 K) = 131.913 J/mol-K – p^o = 0.1 MPa



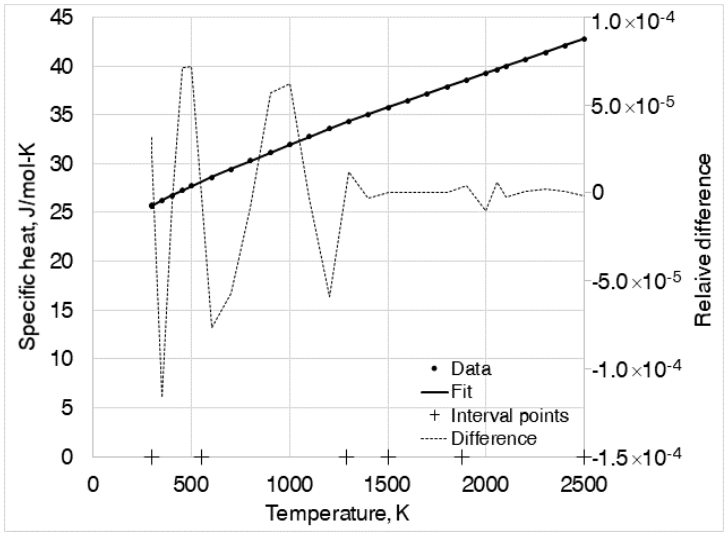


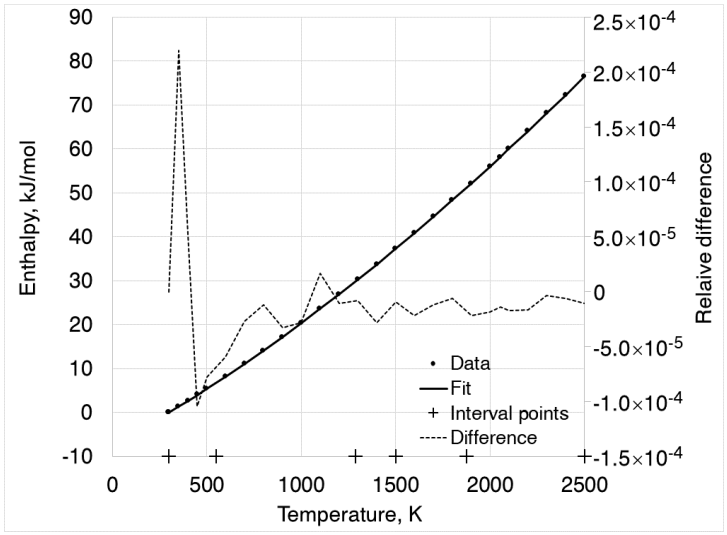


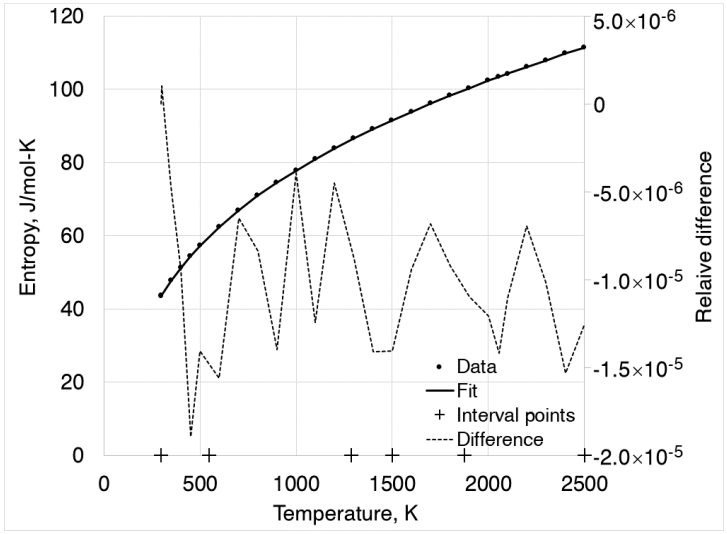


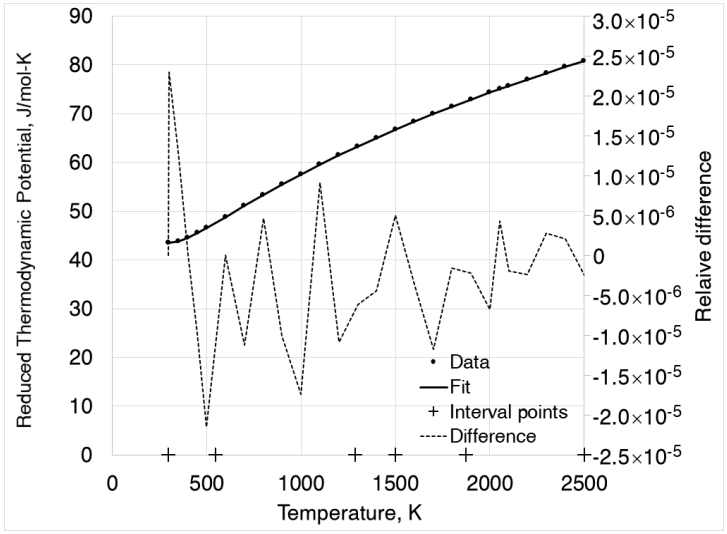
Hafnium, Alpha (Hf) – Hf1(cr)_01

MW = 178.49 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 43.56 J/mol-K – p^o = 0.1 MPa



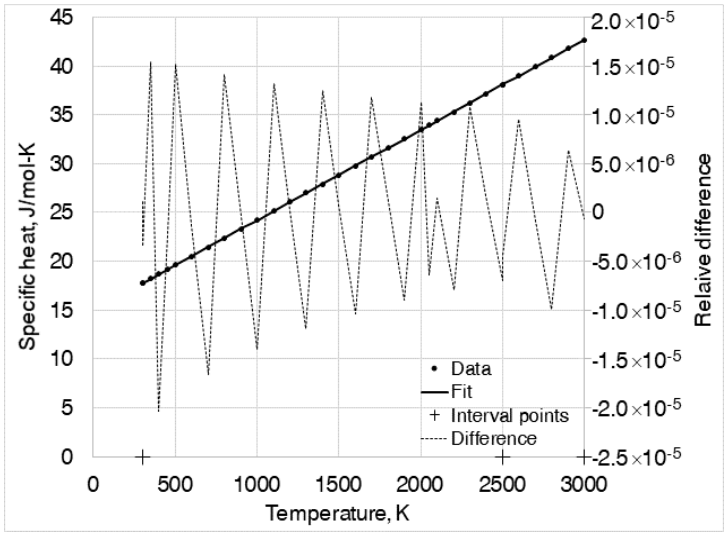


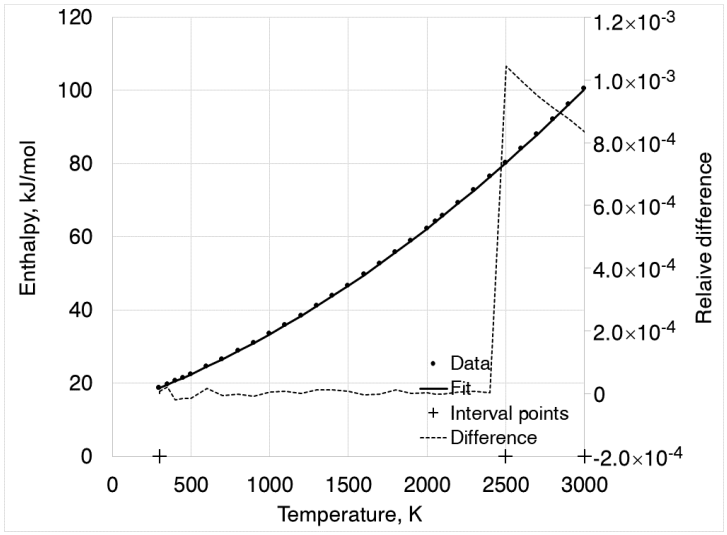


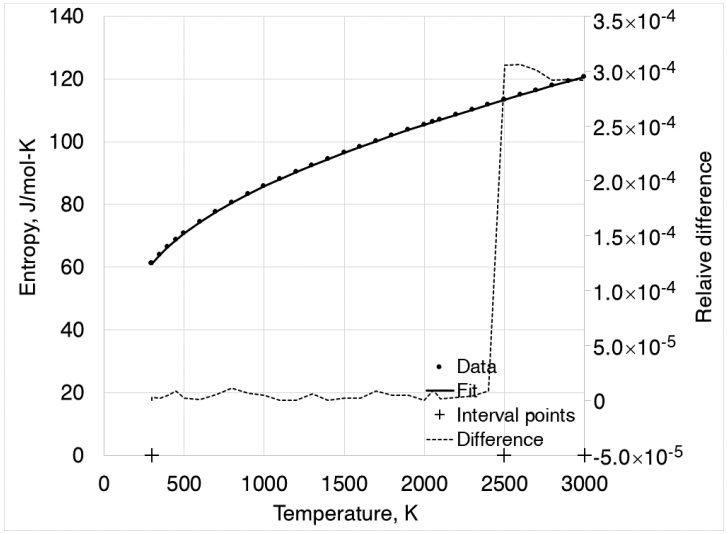


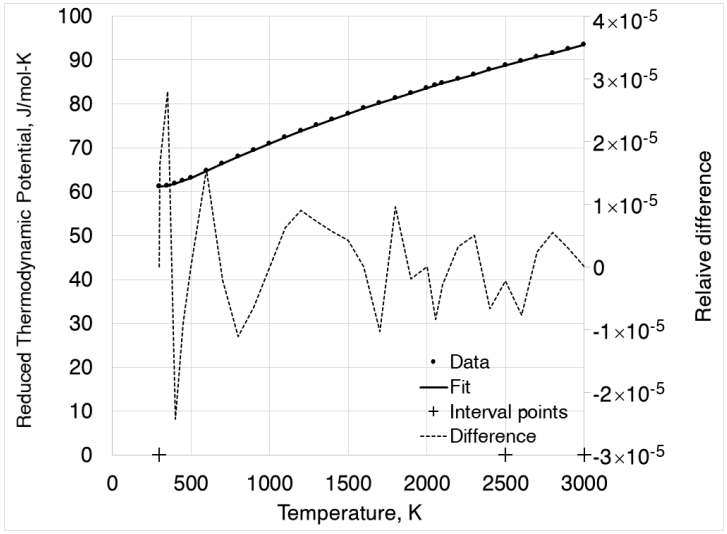
Hafnium, Beta (Hf) – Hf1(cr)_02

MW = 178.49 g/mol – $\Delta_f H^o$(298.15 K) = 18.686 kJ/mol – S^o(298.15 K) = 61.104 J/mol-K – p^o = 0.1 MPa



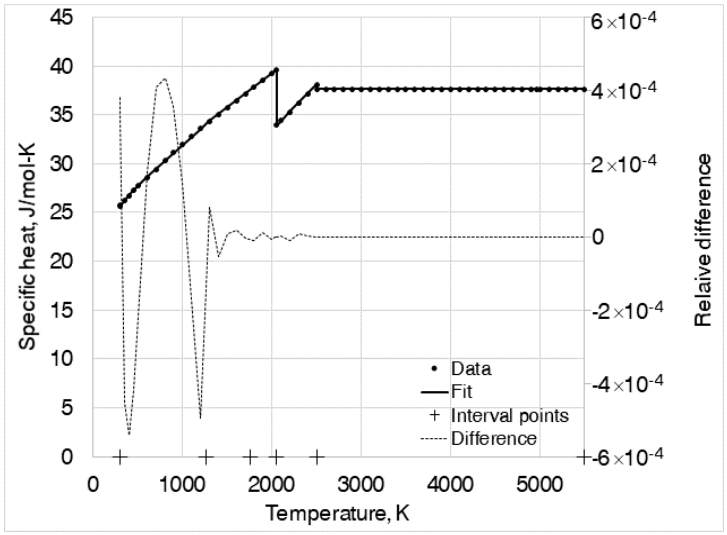


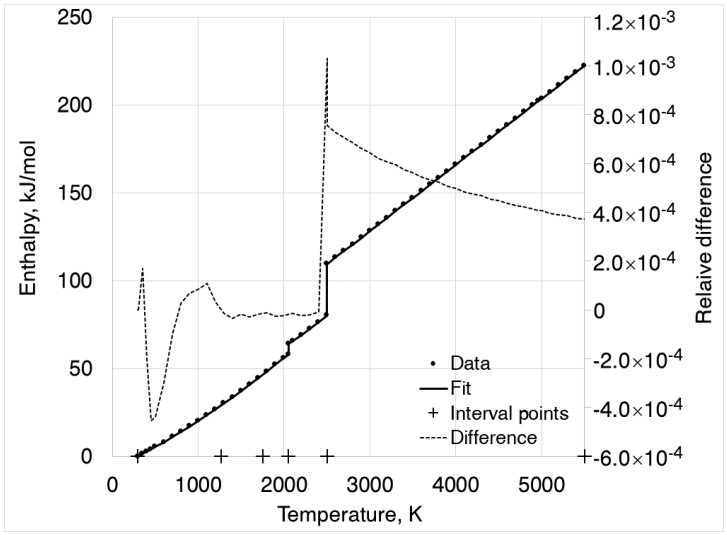


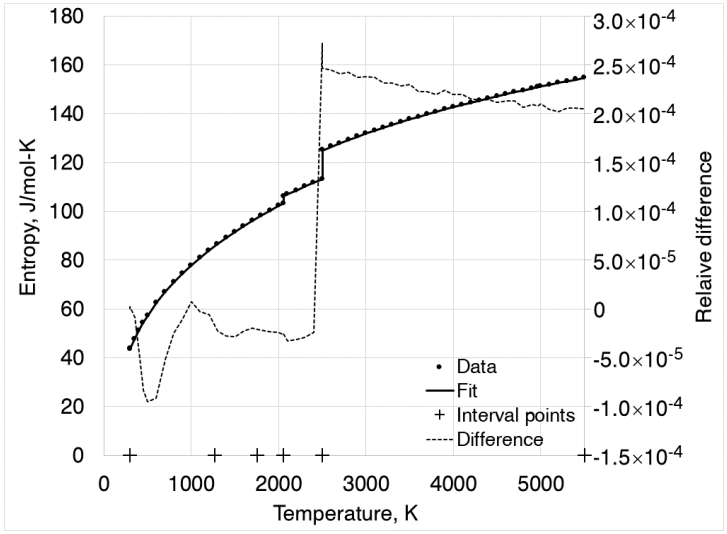


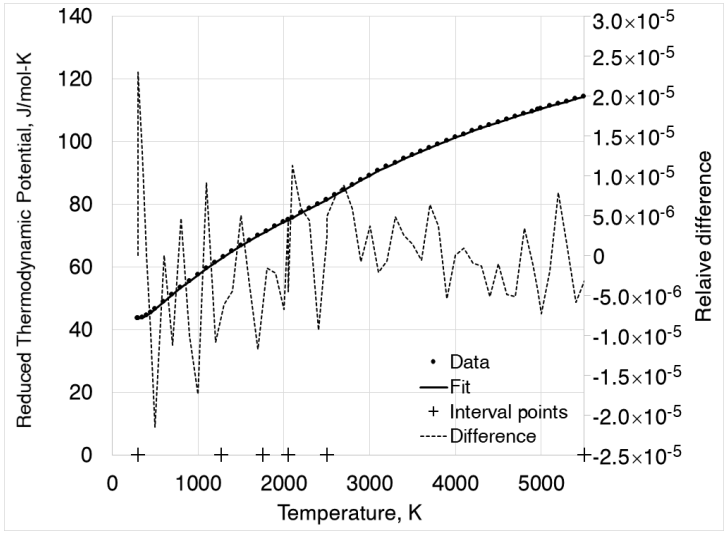
Hafnium (Hf) – Hf1(cr,l)

MW = 178.49 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 43.56 J/mol-K – p^o = 0.1 MPa



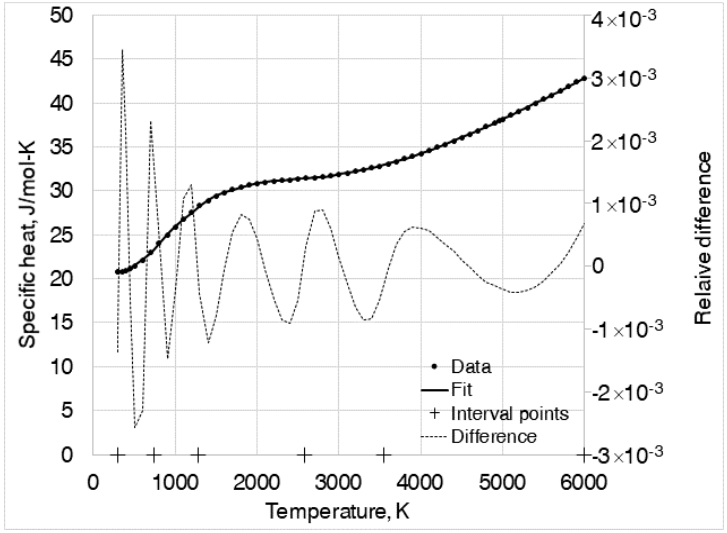


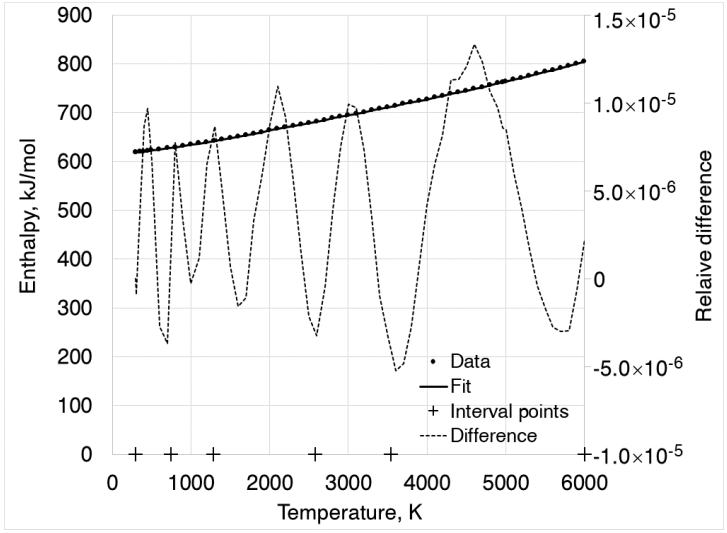


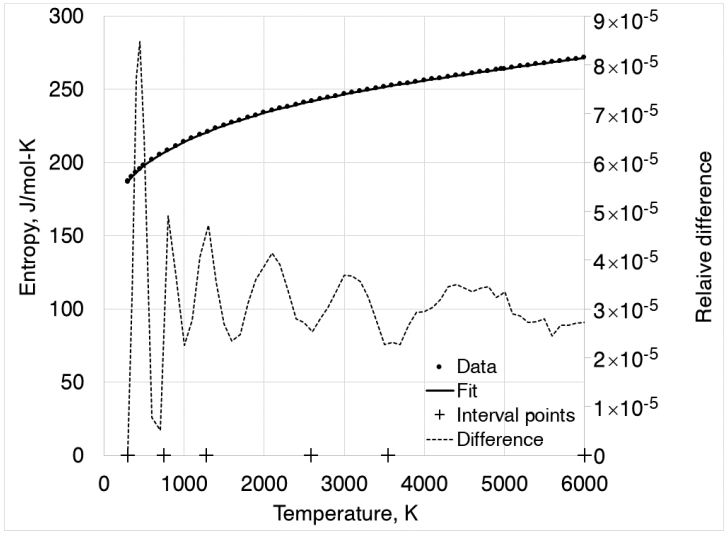


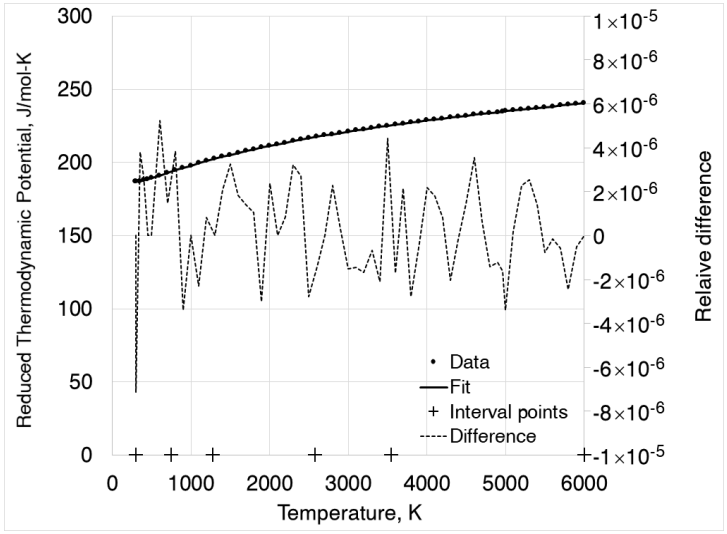
Hafnium (Hf) – Hf1(g)

MW = 178.49 g/mol – $\Delta_f H^o$(298.15 K) = 618.395 kJ/mol – S^o(298.15 K) = 186.897 J/mol-K – p^o = 0.1 MPa



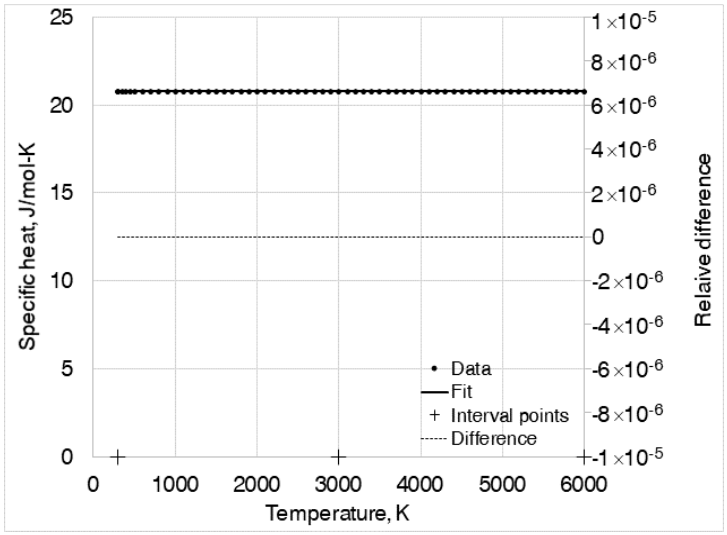


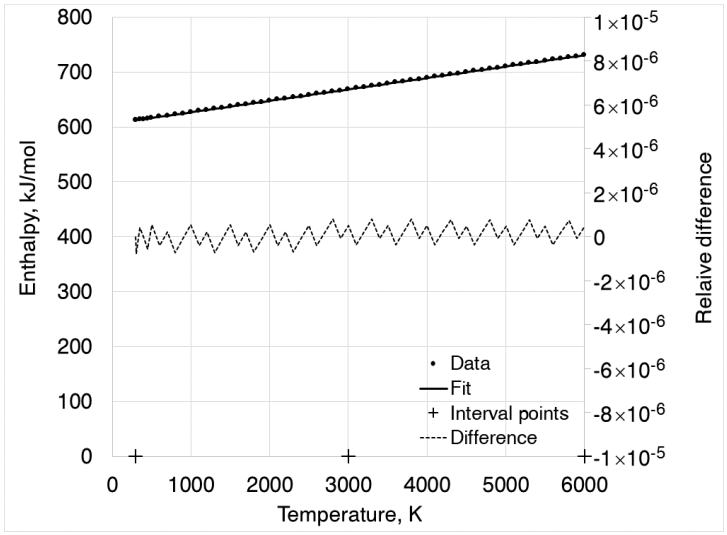


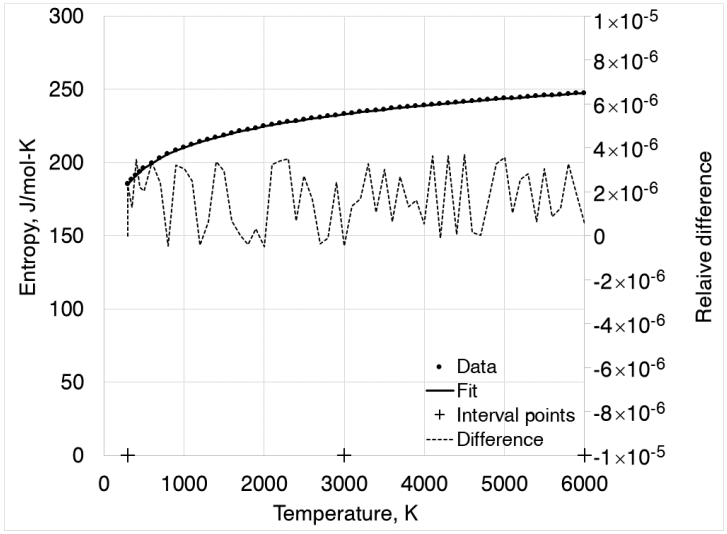


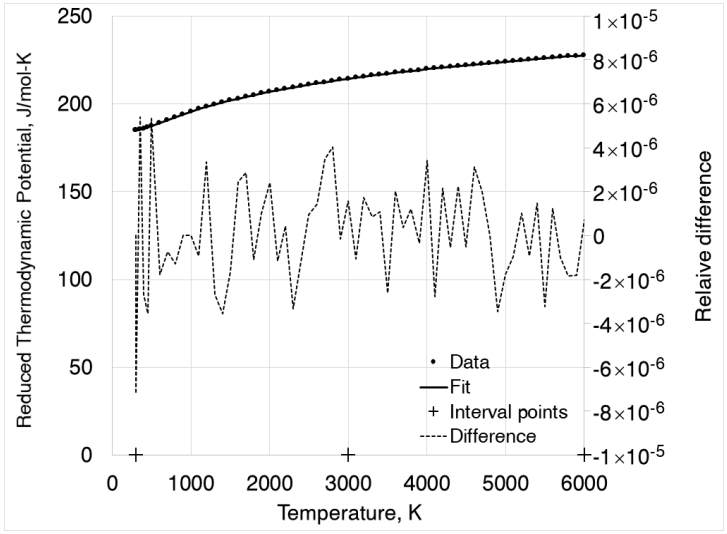
Hafnium, Ion (Hf^-) – Hf1-(g)

MW = 178.49 g/mol – $\Delta_f H^o$(298.15 K) = 612.197 kJ/mol – S^o(298.15 K) = 185.04 J/mol-K – p^o = 0.1 MPa



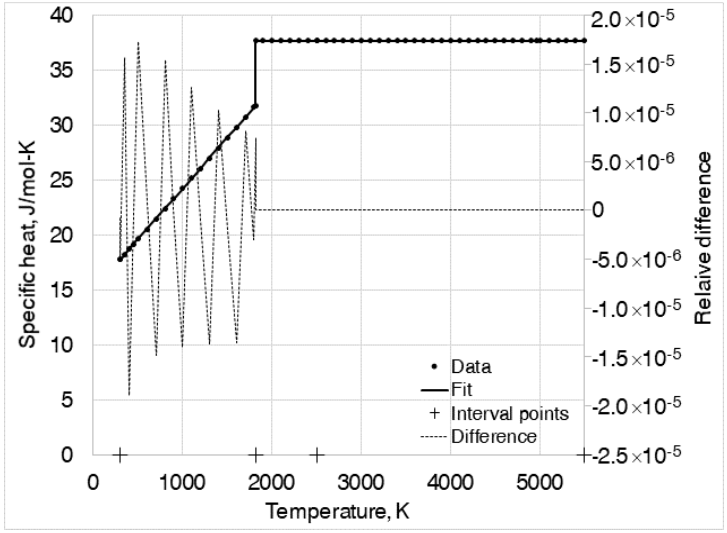


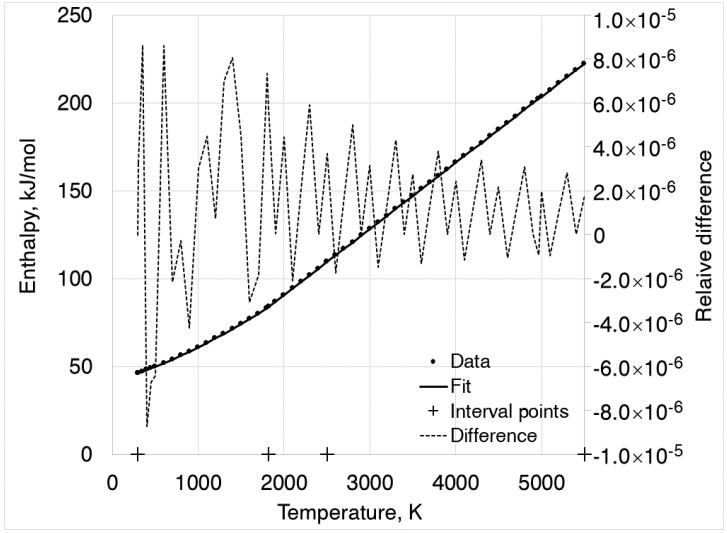


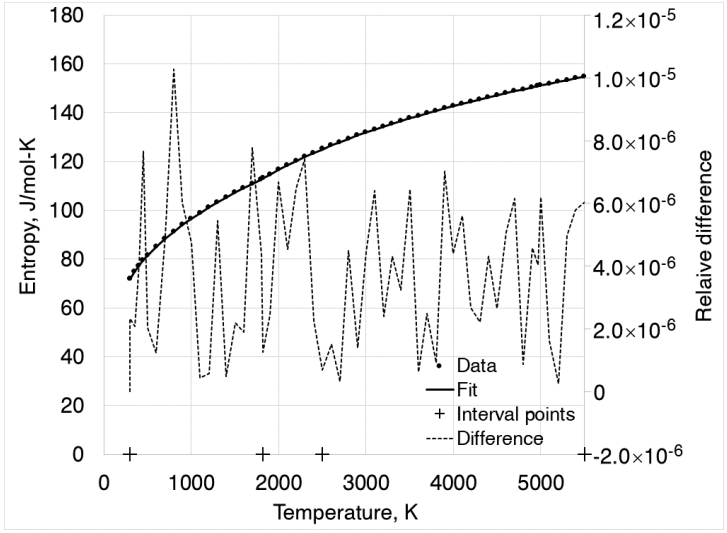


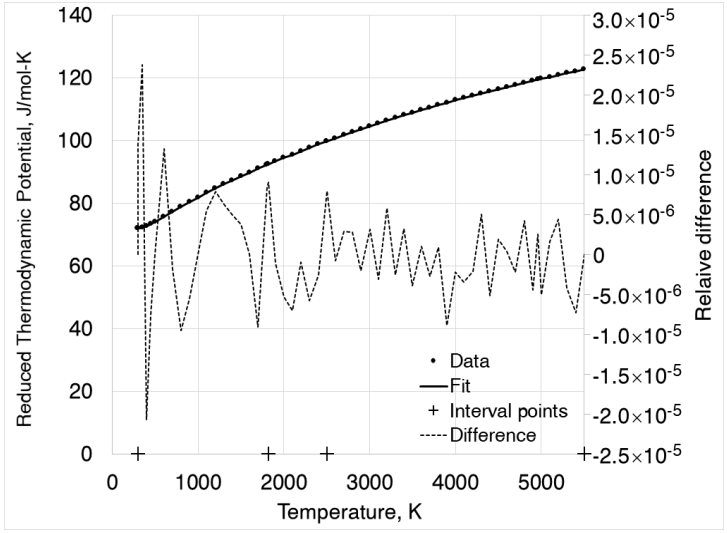
Hafnium (Hf) – Hf1(l)

MW = 178.49 g/mol – $\Delta_f H^o$(298.15 K) = 46.209 kJ/mol – S^o(298.15 K) = 71.938 J/mol-K – p^o = 0.1 MPa



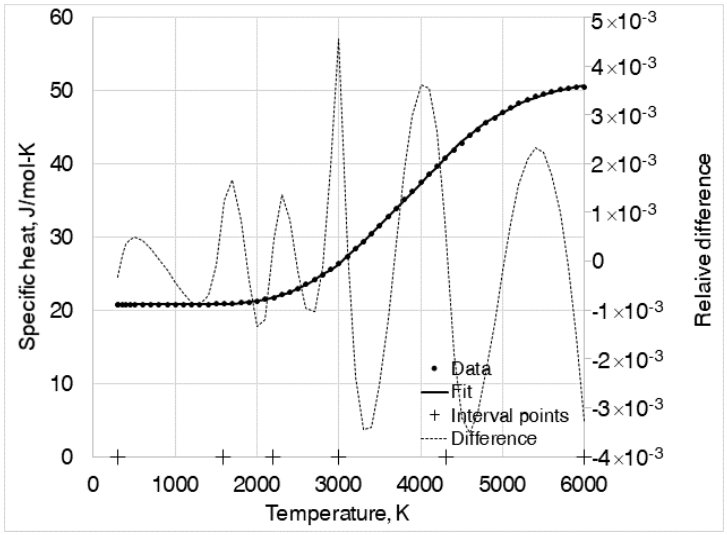


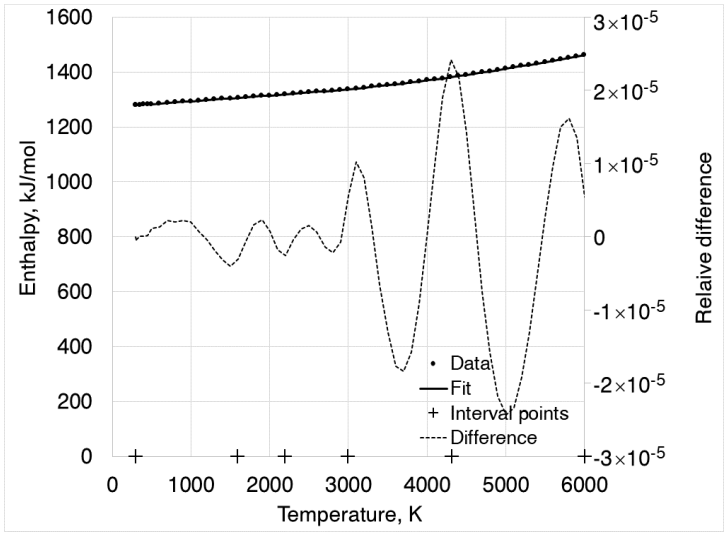


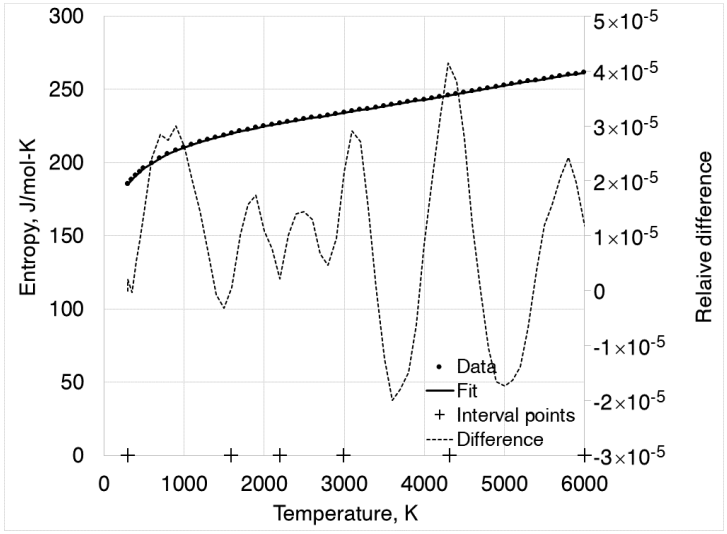


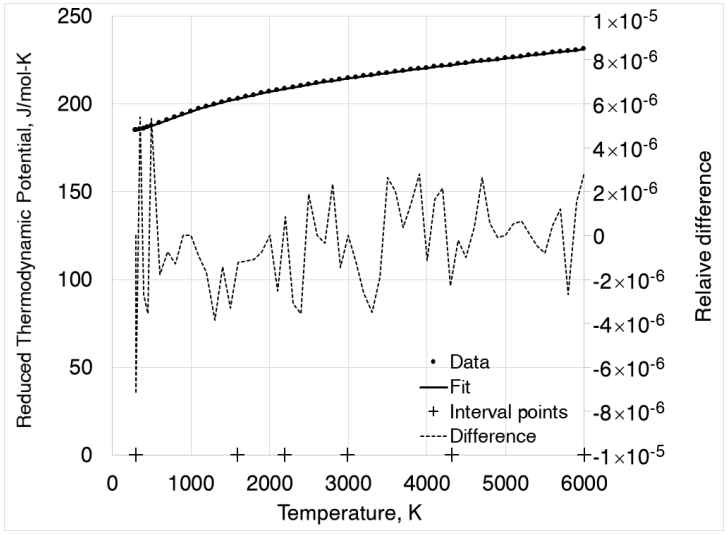
Hafnium, Ion (Hf^+) – Hf1+(g)

MW = 178.49 g/mol – $\Delta_f H^o$(298.15 K) = 1278.95 kJ/mol – S^o(298.15 K) = 185.04 J/mol-K – p^o = 0.1 MPa



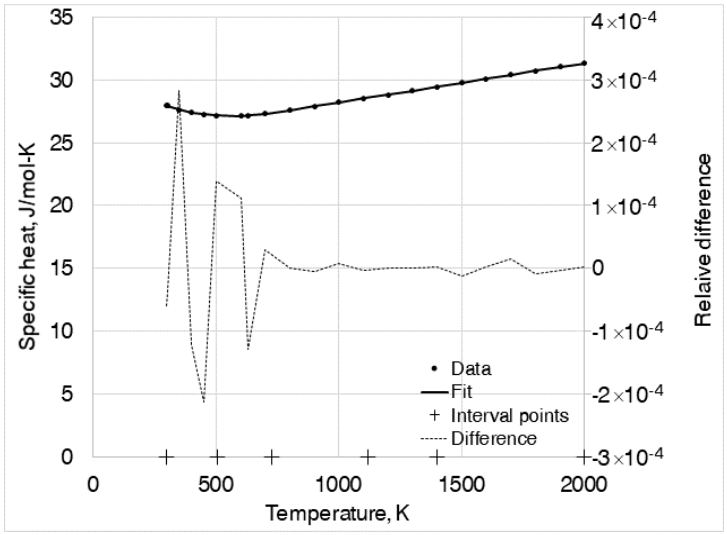


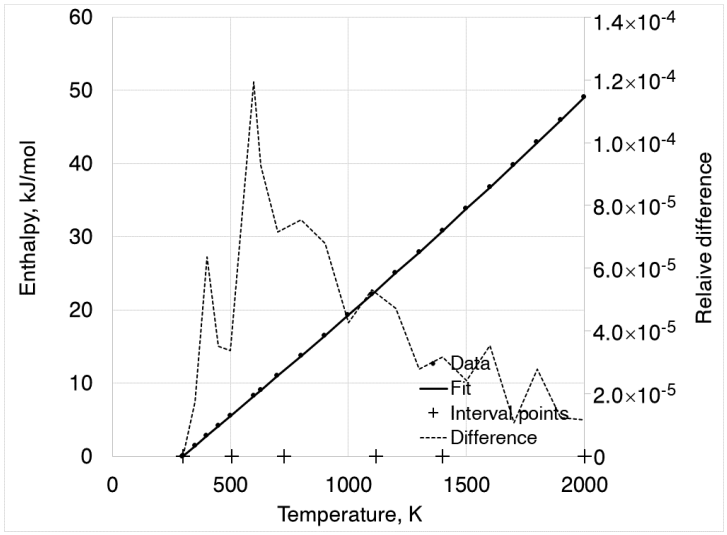


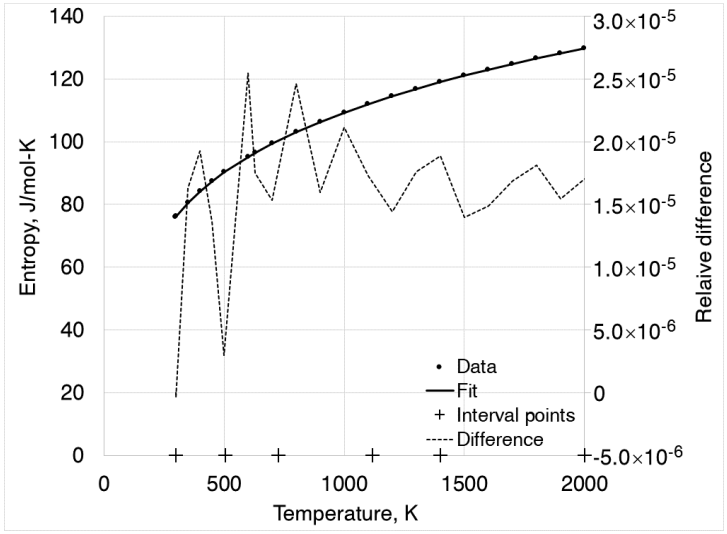


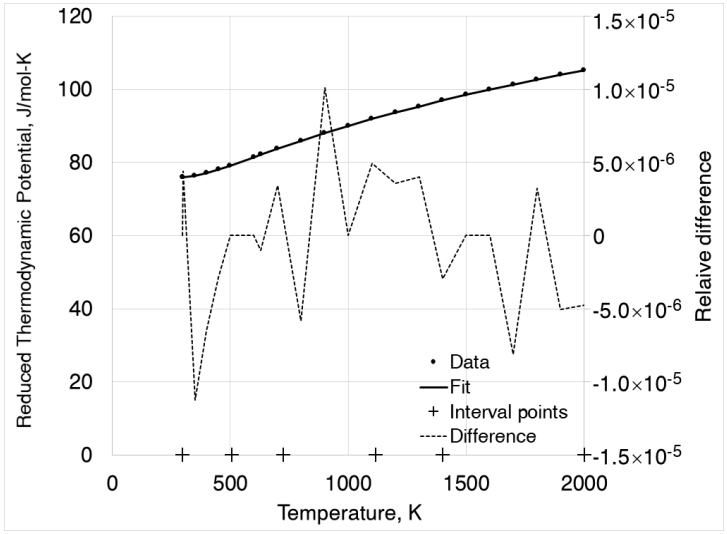
Mercury (Hg) – Hg1(cr,l)

MW = 200.59 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 76.028 J/mol-K – p^o = 0.1 MPa



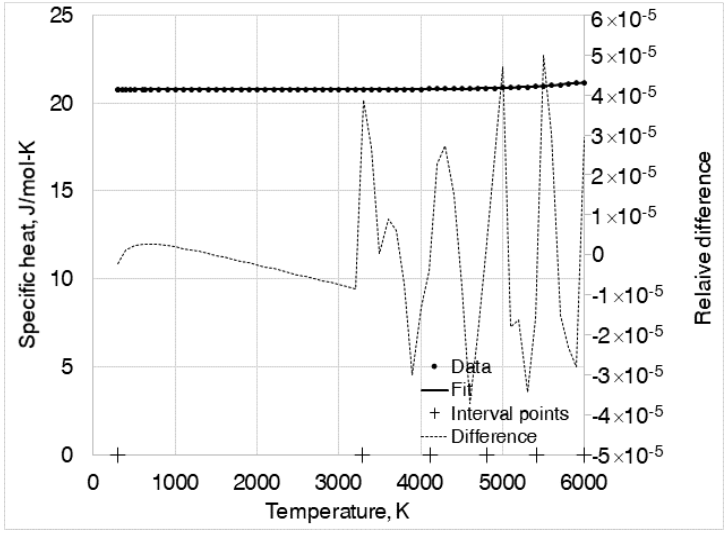


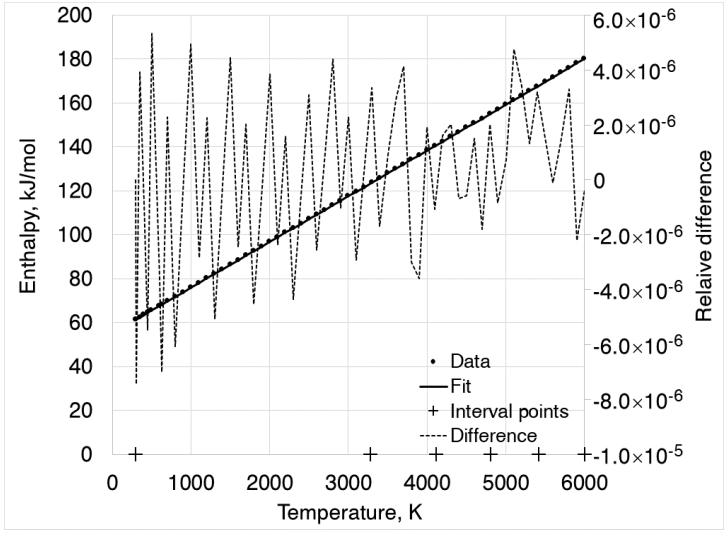


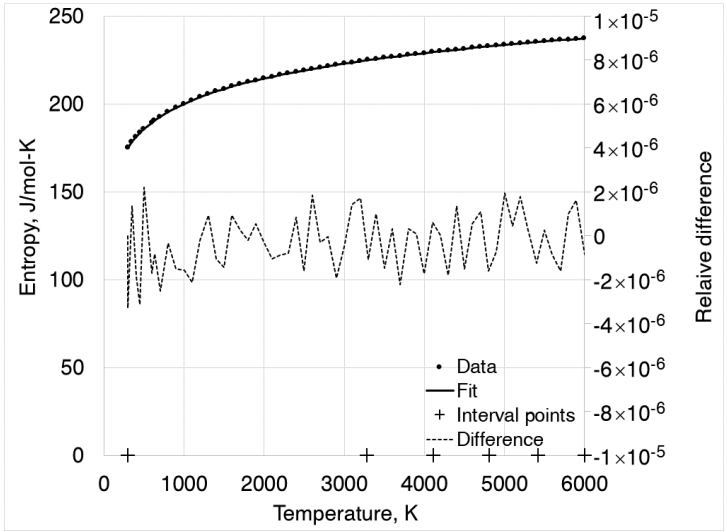


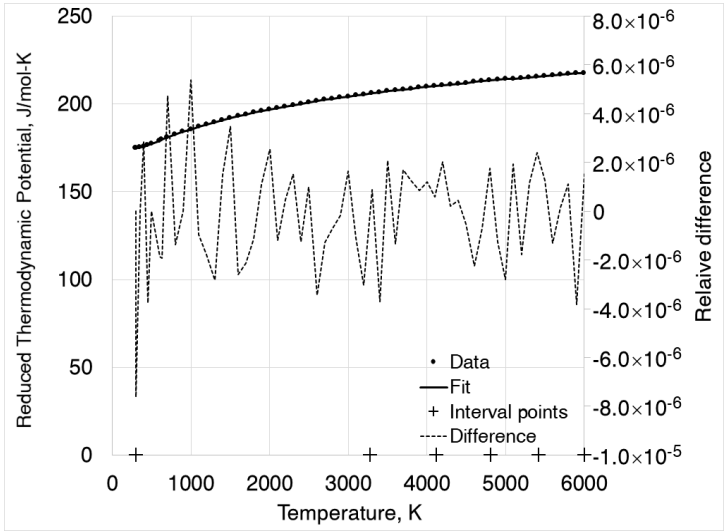
Mercury (Hg) – Hg1(g)

MW = 200.59 g/mol – $\Delta_f H^o$(298.15 K) = 61.38 kJ/mol – S^o(298.15 K) = 174.97 J/mol-K – p^o = 0.1 MPa



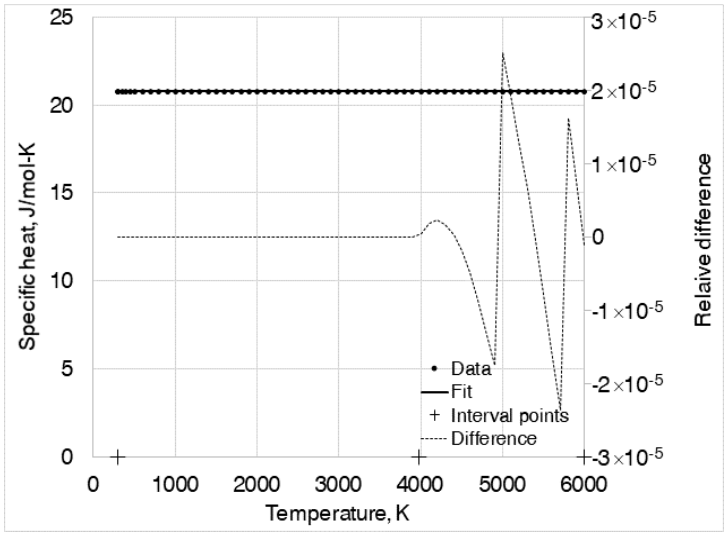


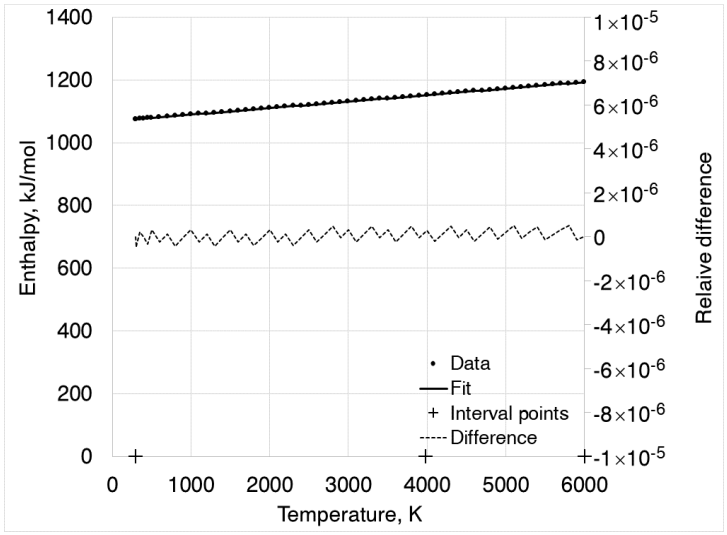


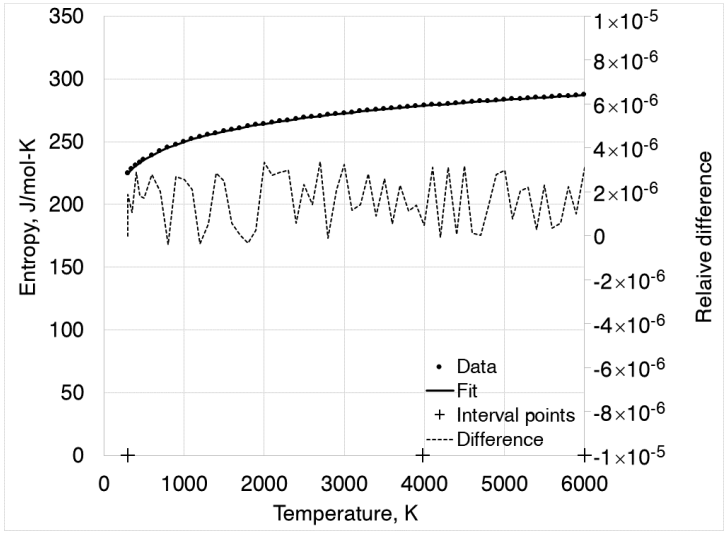


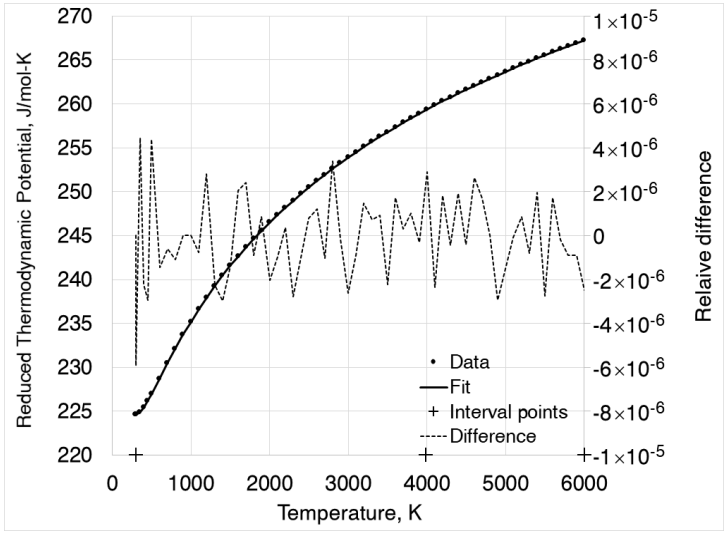
Mercury, Ion (Hg^+) – Hg1+(g)

MW = 200.59 g/mol – $\Delta_f H^o$(298.15 K) = 1074.623 kJ/mol – S^o(298.15 K) = 224.617 J/mol-K – p^o = 0.1 MPa



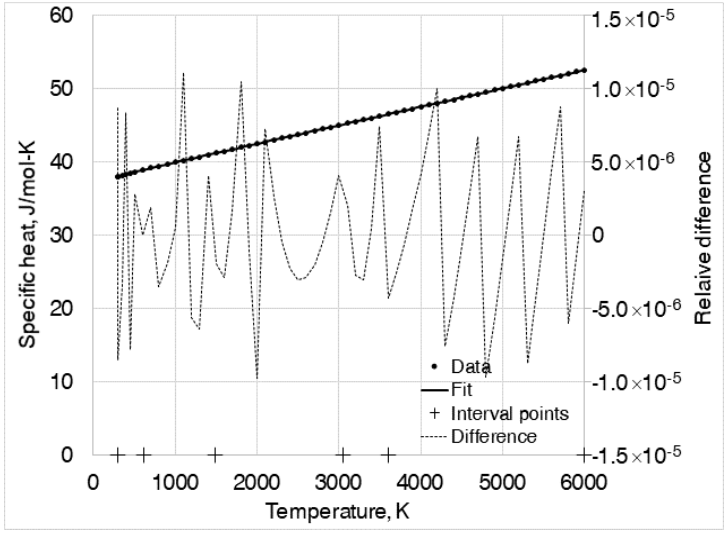


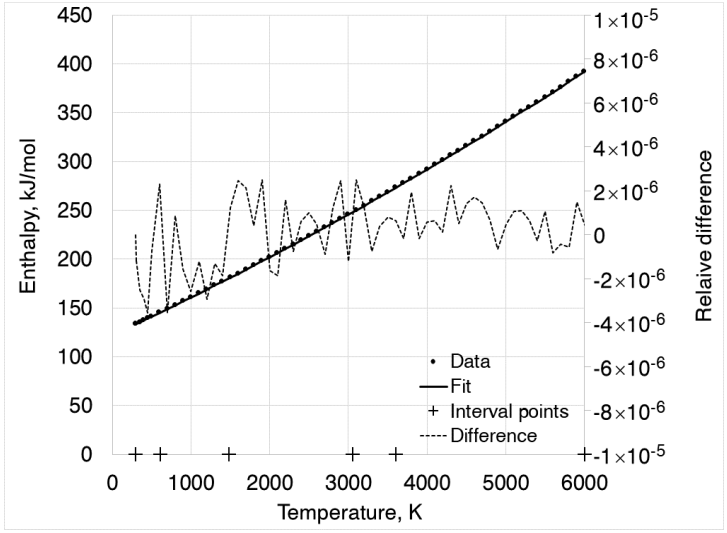


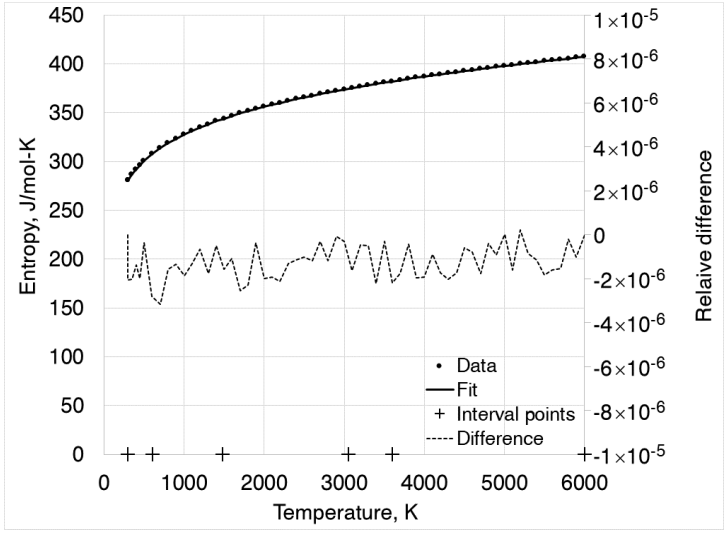


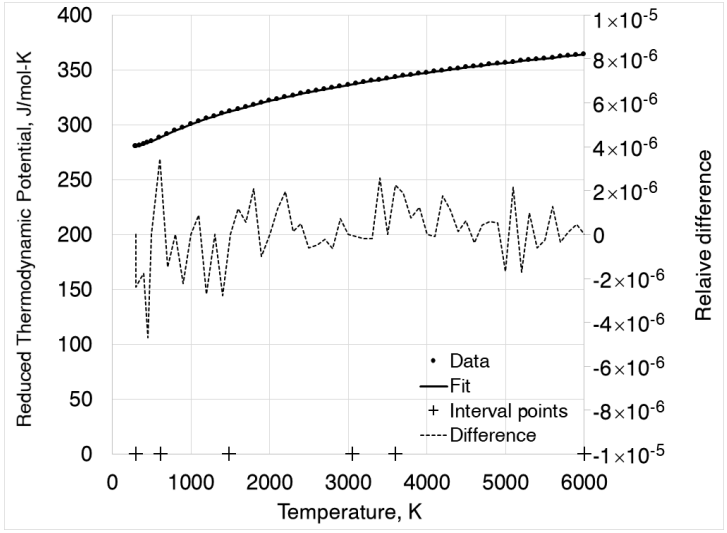
Mercury Iodide (HgI) – Hg1I1(g)

MW = 327.49 g/mol – $\Delta_f H^o$(298.15 K) = 133.47 kJ/mol – S^o(298.15 K) = 280.752 J/mol-K – p^o = 0.1 MPa



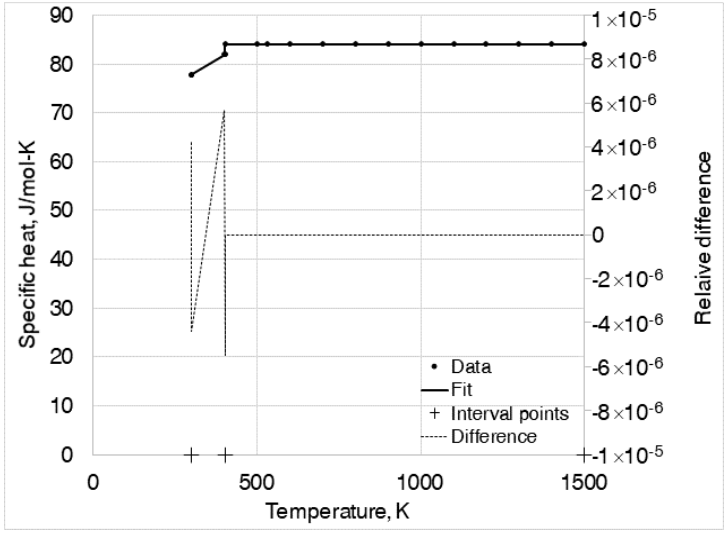


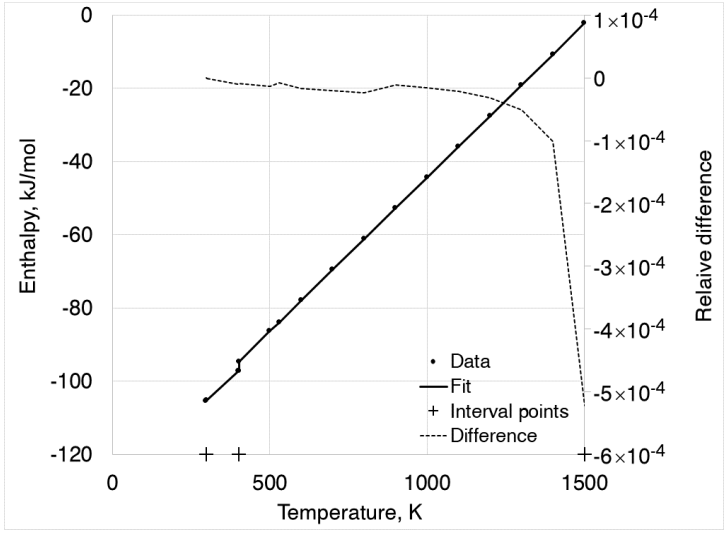


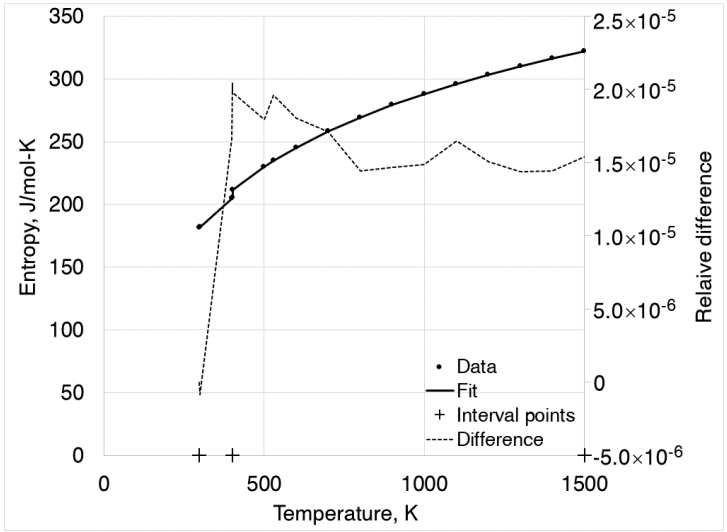


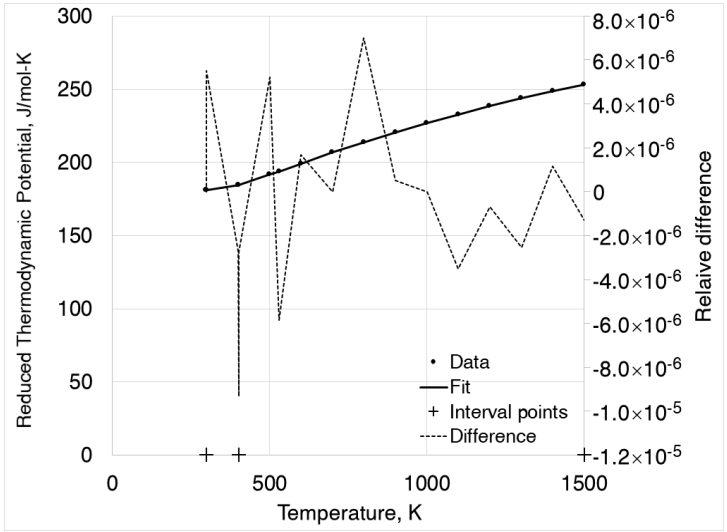
Mercury Iodide (HgI_2) – Hg1I2(cr)

MW = 454.4 g/mol – $\Delta_f H^o$(298.15 K) = -105.437 kJ/mol – S^o(298.15 K) = 181.326 J/mol-K – p^o = 0.1 MPa



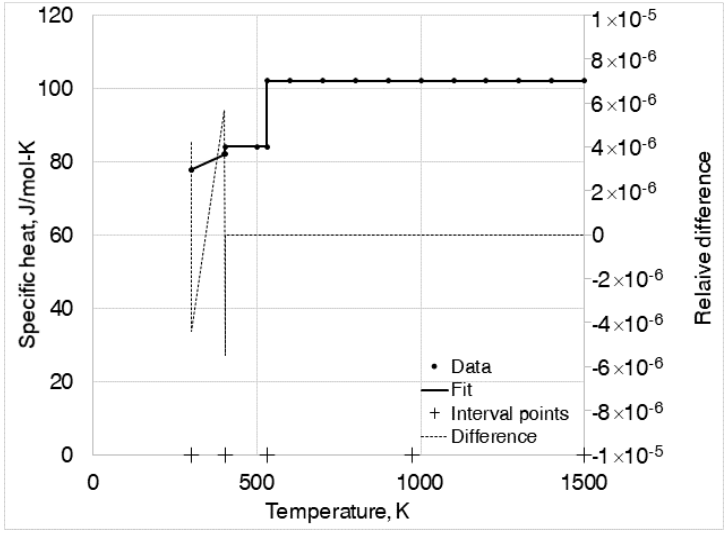


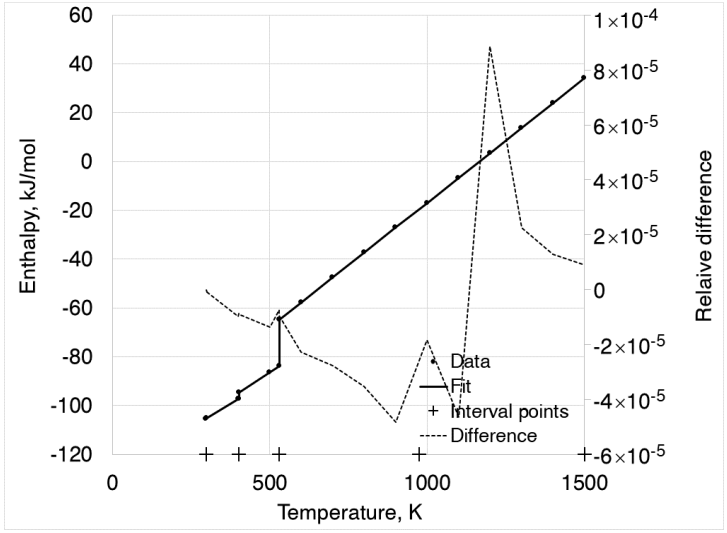


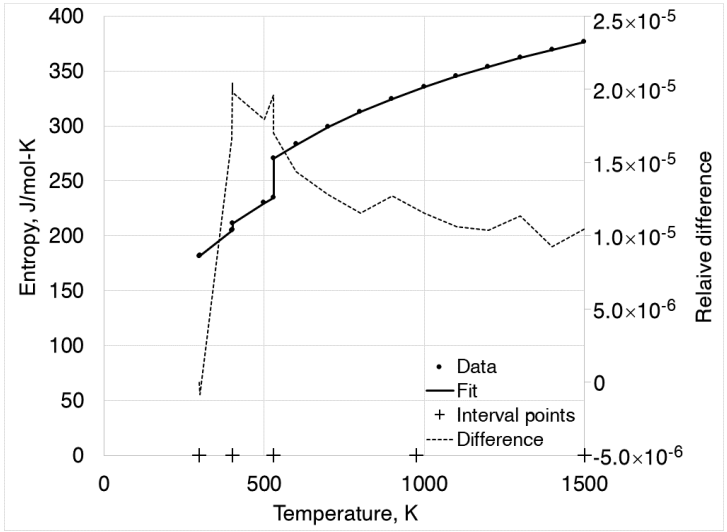


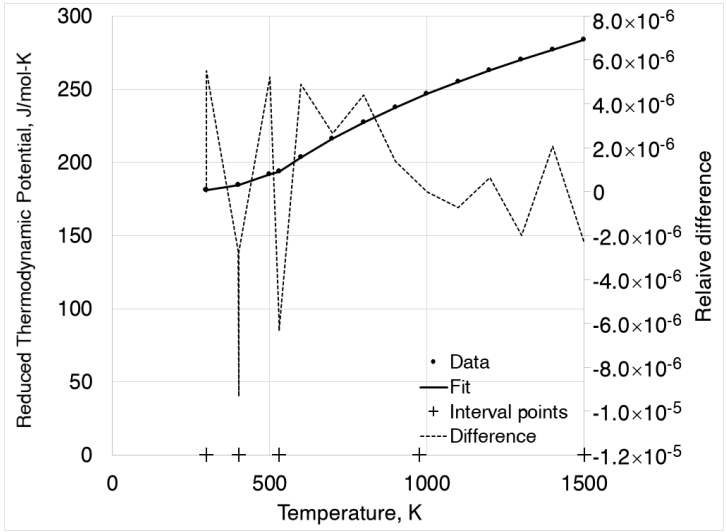
Mercury Iodide (HgI_2) – Hg1I2(cr,l)

MW = 454.4 g/mol – $\Delta_f H^o$(298.15 K) = -105.437 kJ/mol – S^o(298.15 K) = 181.326 J/mol-K – p^o = 0.1 MPa



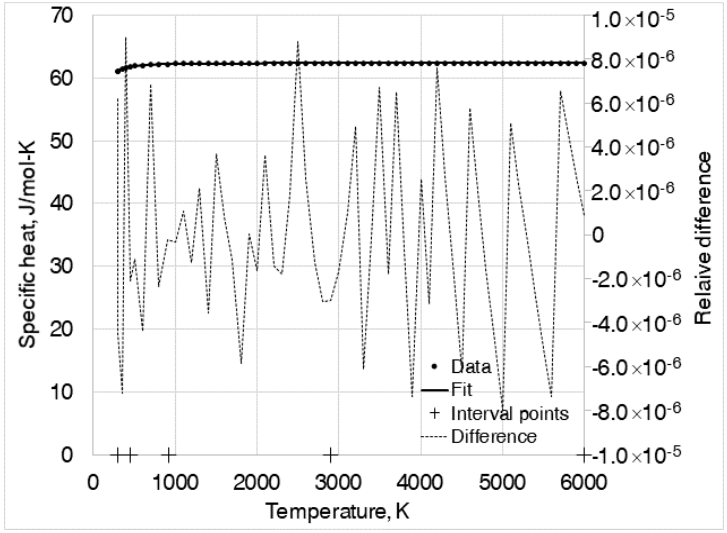


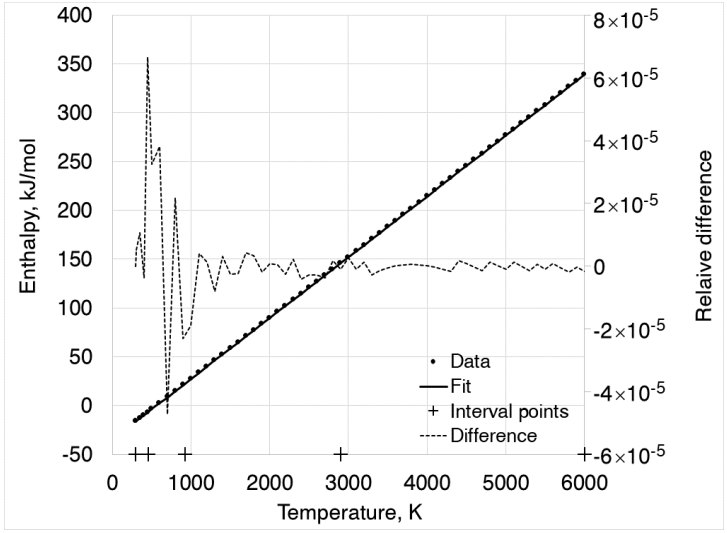


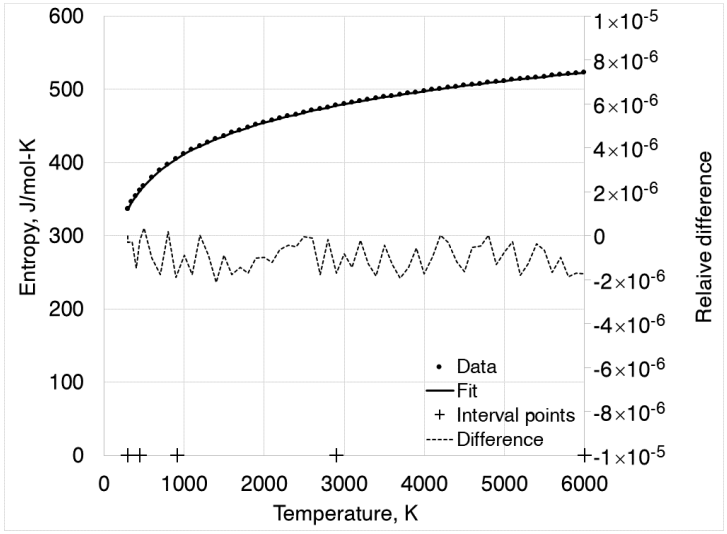


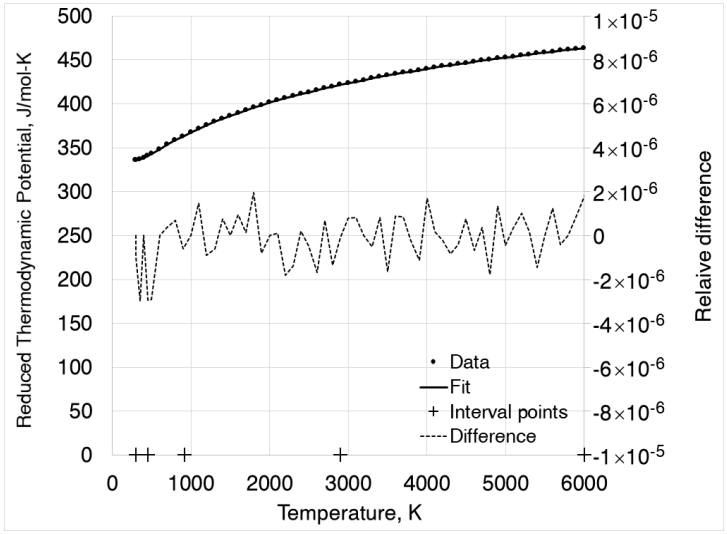
Mercury Iodide (HgI_2) – Hg1I2(g)

MW = 454.4 g/mol – $\Delta_f H^o$(298.15 K) = -16.129 kJ/mol – S^o(298.15 K) = 336.206 J/mol-K – p^o = 0.1 MPa



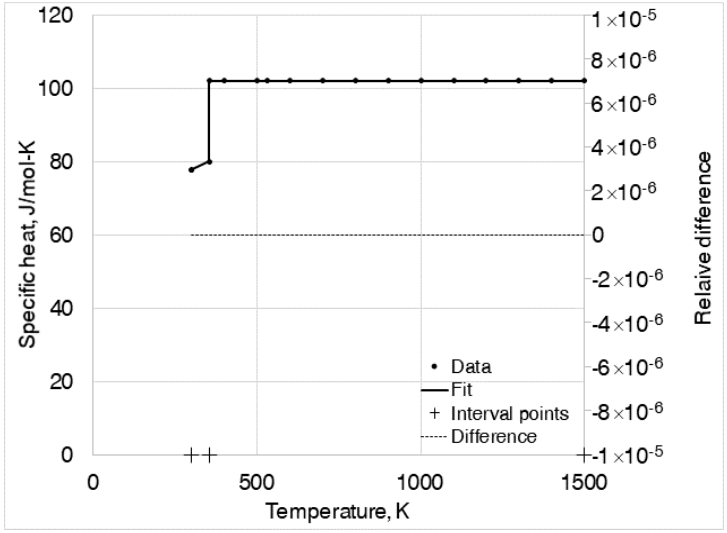


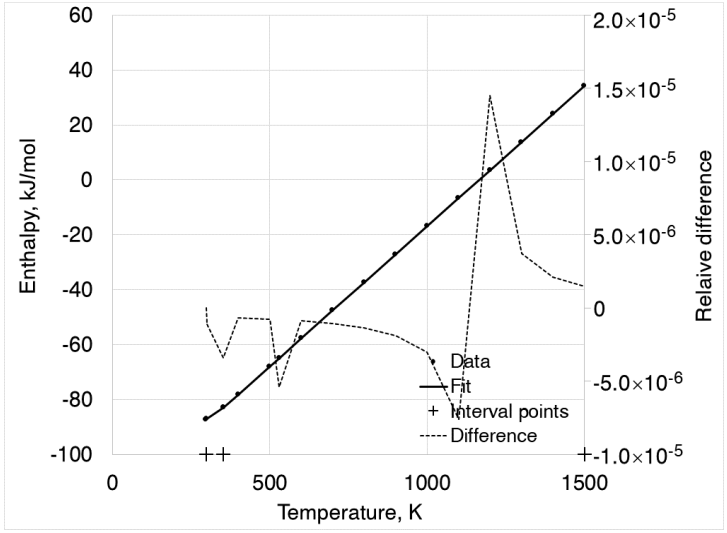


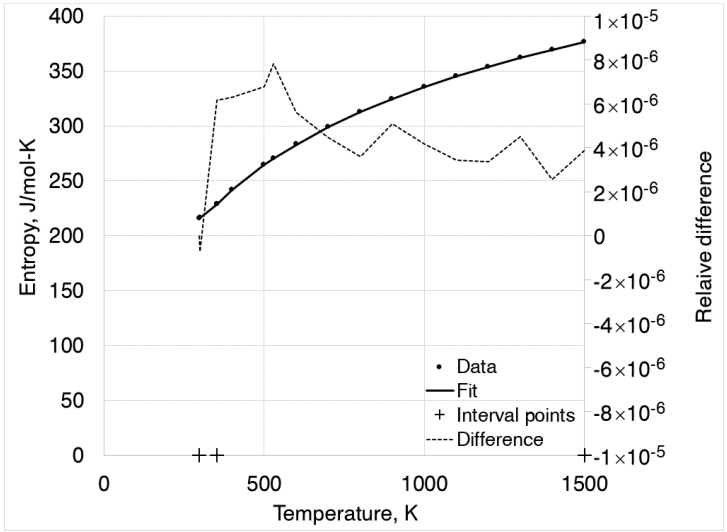


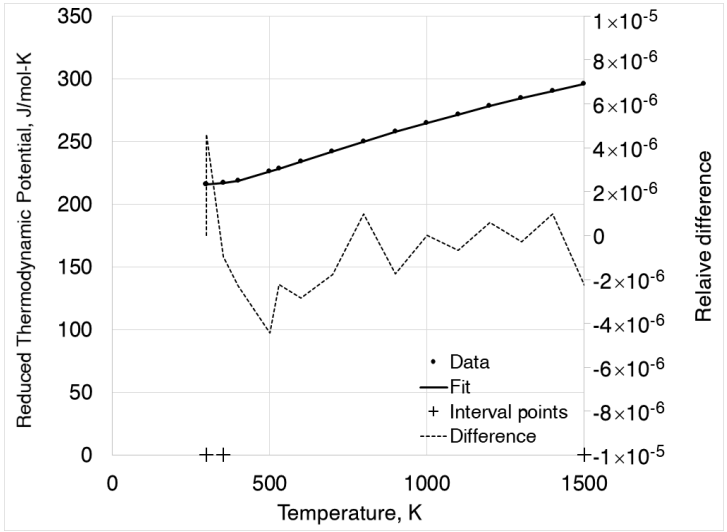
Mercury Iodide (HgI_2) – Hg1I2(l)

MW = 454.4 g/mol – $\Delta_f H^o$(298.15 K) = -87.287 kJ/mol – S^o(298.15 K) = 215.663 J/mol-K – p^o = 0.1 MPa



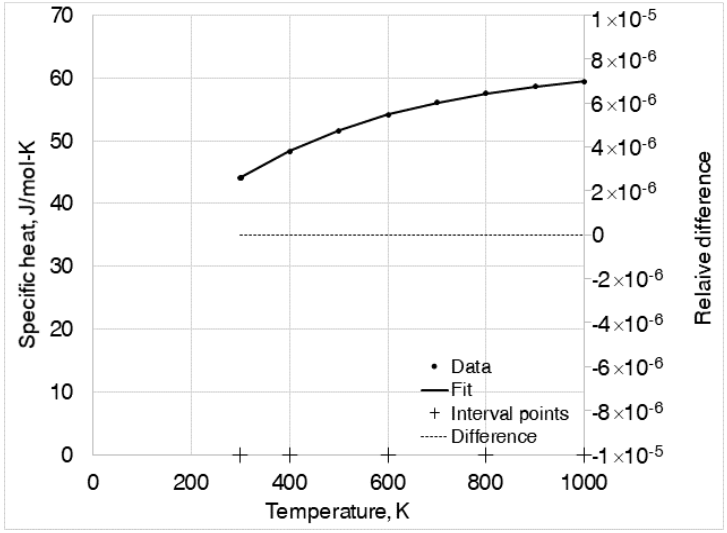


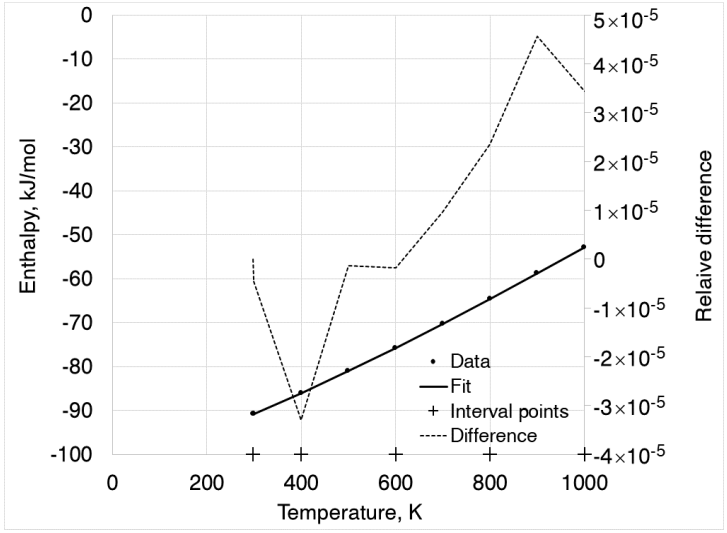


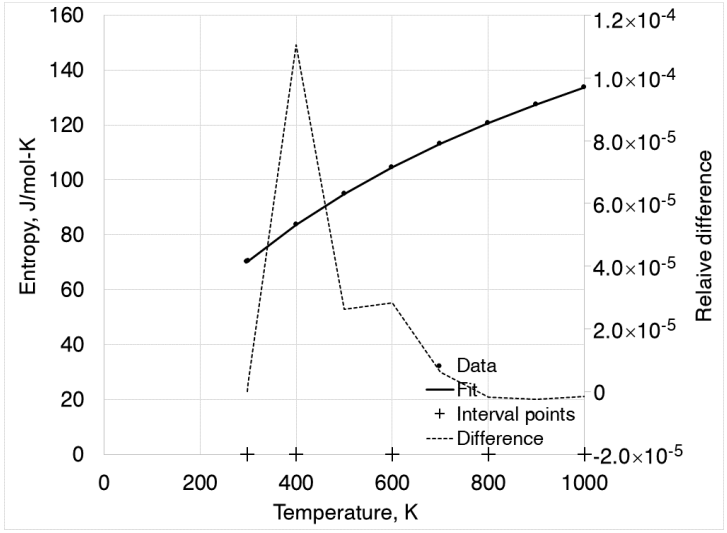


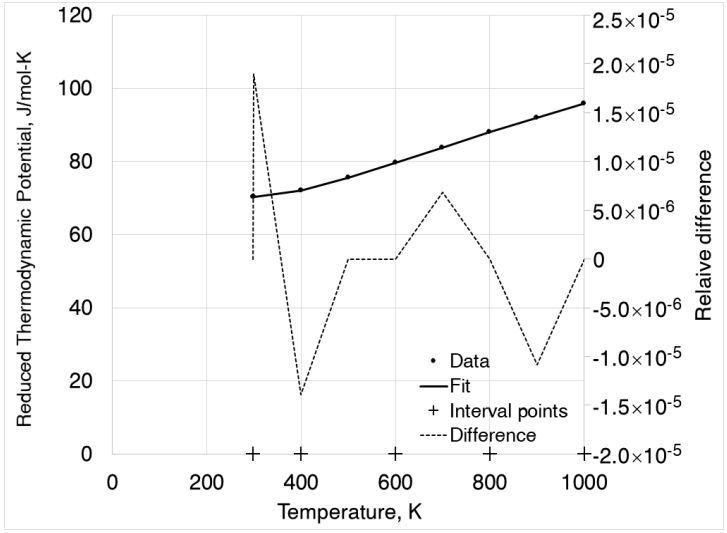
Mercury Oxide (HgO) – Hg1O1(cr)

MW = 216.59 g/mol – $\Delta_f H^o$(298.15 K) = -90.789 kJ/mol – S^o(298.15 K) = 70.27 J/mol-K – p^o = 0.1 MPa



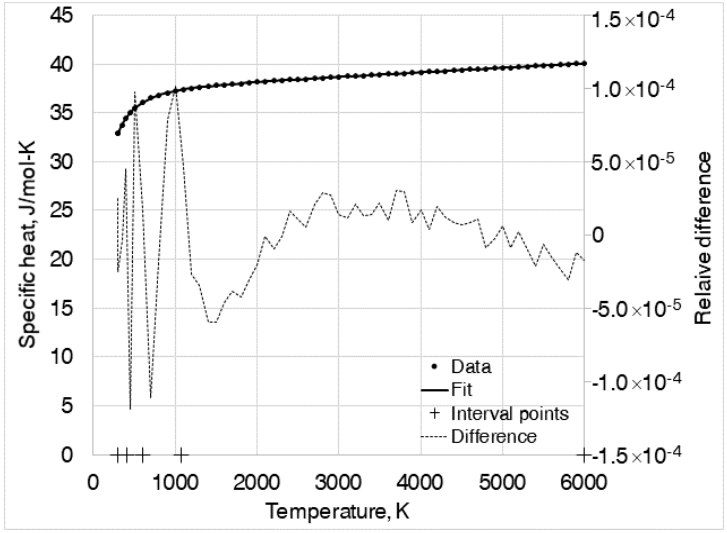


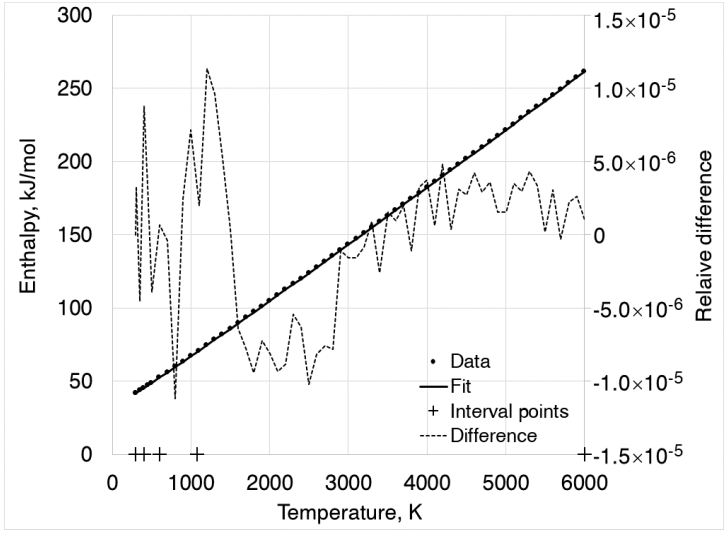


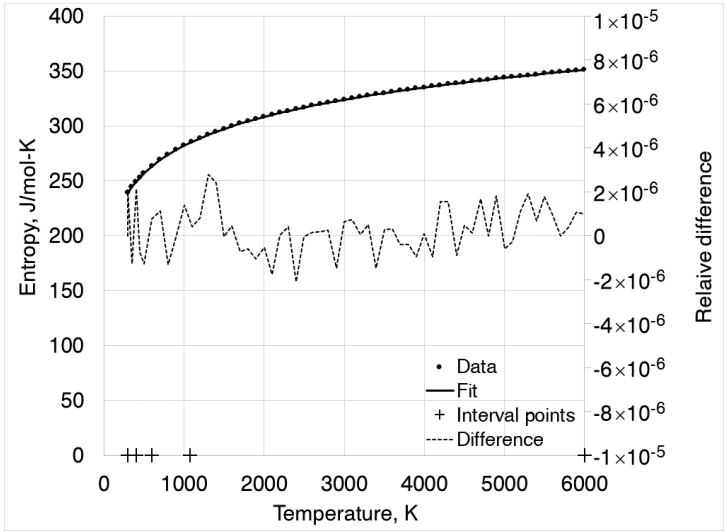


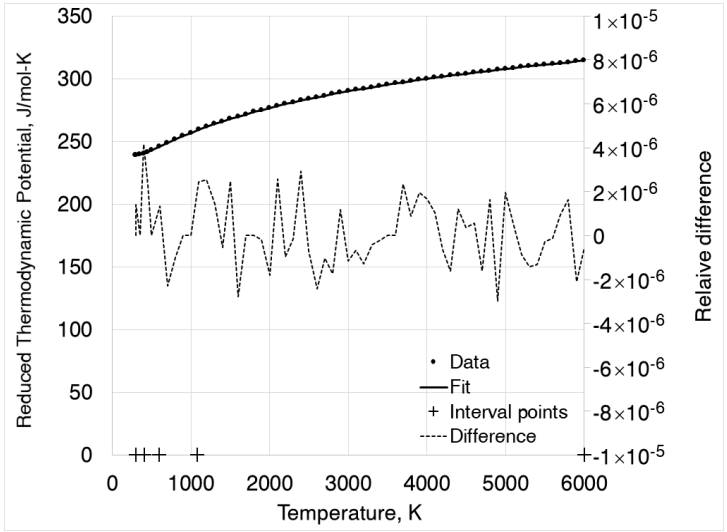
Mercury Oxide (HgO) – Hg1O1(g)

MW = 216.59 g/mol – $\Delta_f H^o$(298.15 K) = 41.84 kJ/mol – S^o(298.15 K) = 239.161 J/mol-K – p^o = 0.1 MPa



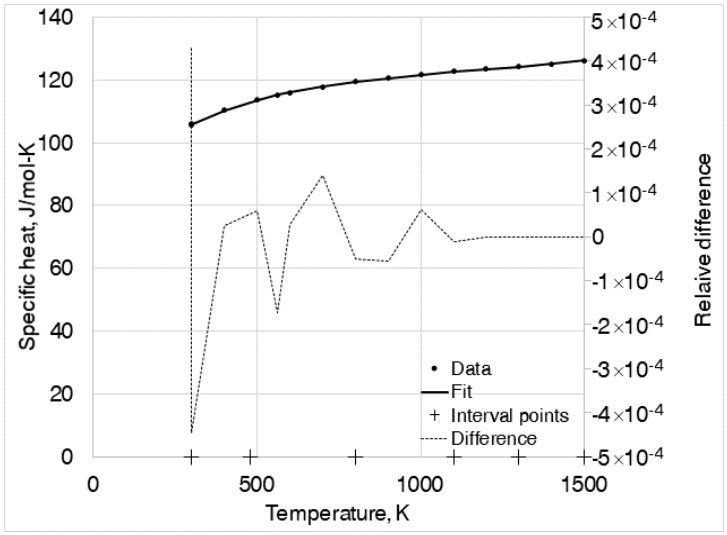


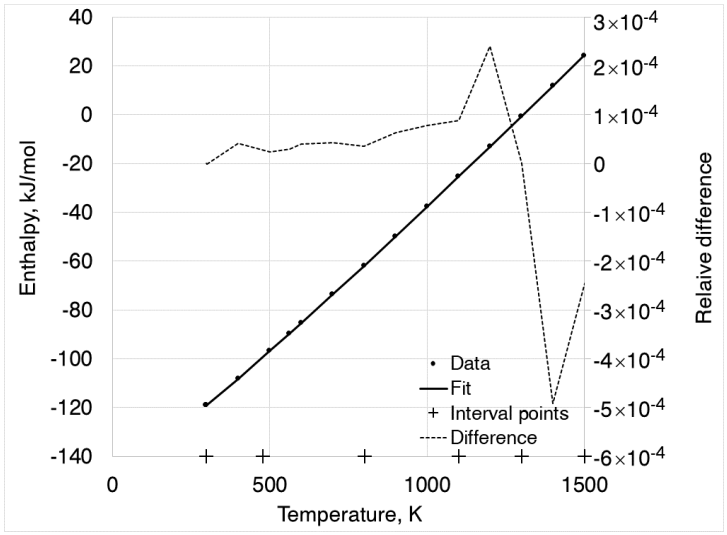


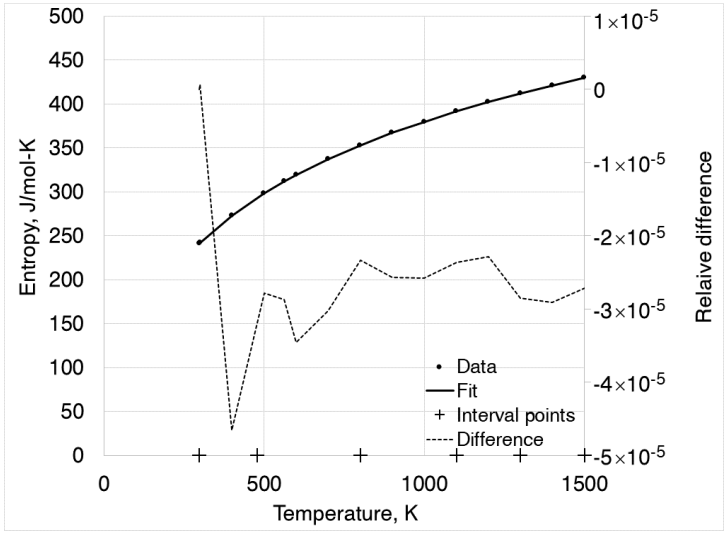


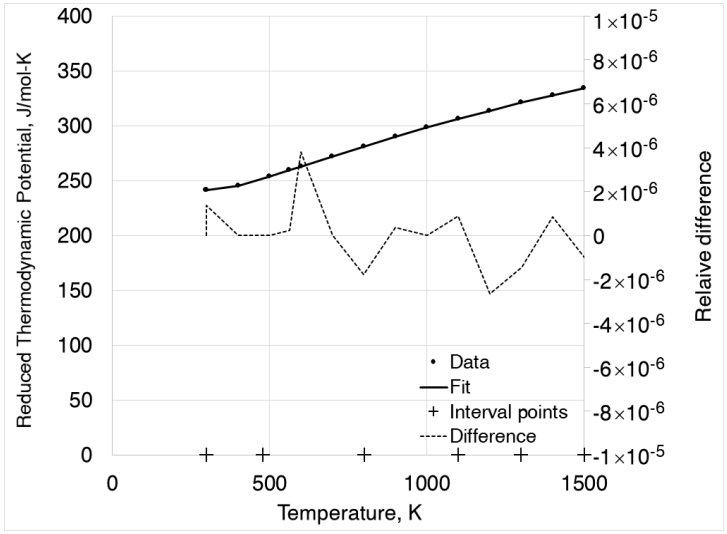
Mercury Iodide (Hg_2I_2) – Hg2I2(cr)

MW = 654.99 g/mol – $\Delta_f H^o$(298.15 K) = -119.085 kJ/mol – S^o(298.15 K) = 241.291 J/mol-K – p^o = 0.1 MPa



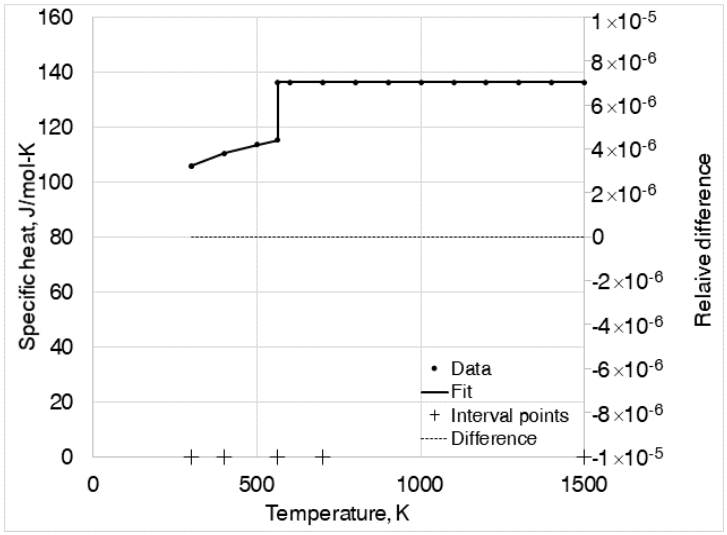


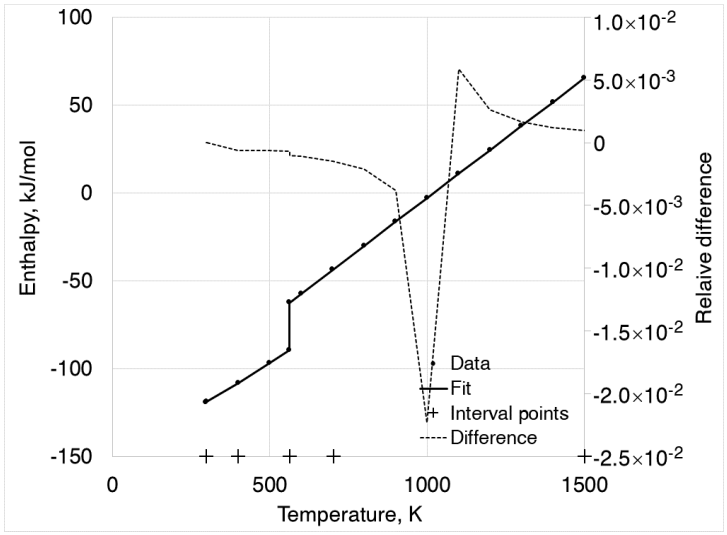


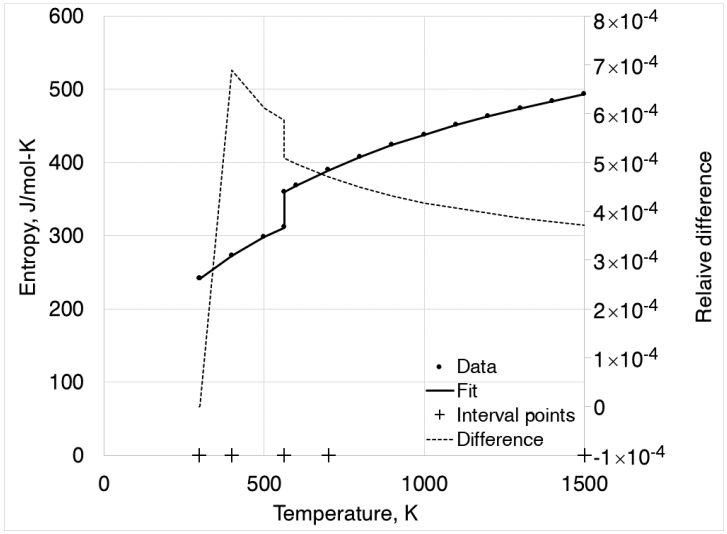


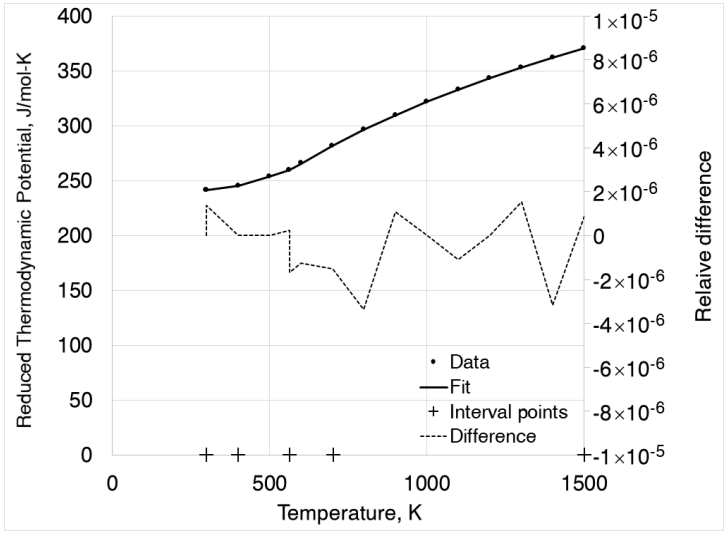
Mercury Iodide (Hg_2I_2) – Hg2I2(cr,l)

MW = 654.99 g/mol – $\Delta_f H^o$(298.15 K) = -119.085 kJ/mol – S^o(298.15 K) = 241.291 J/mol-K – p^o = 0.1 MPa



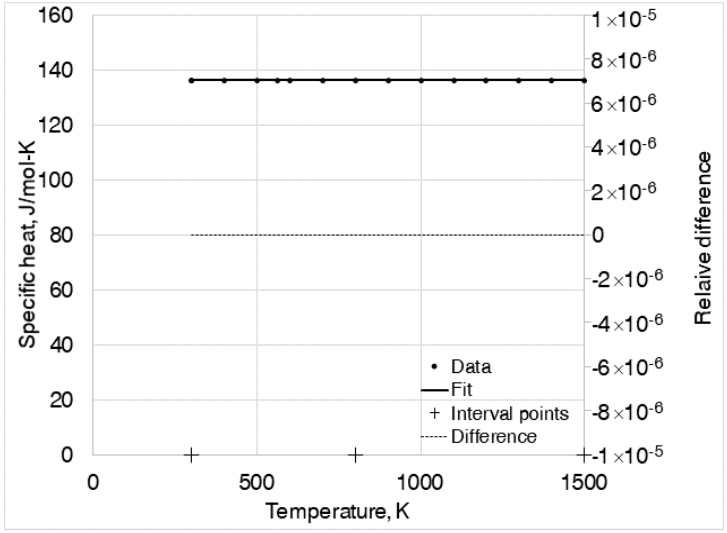


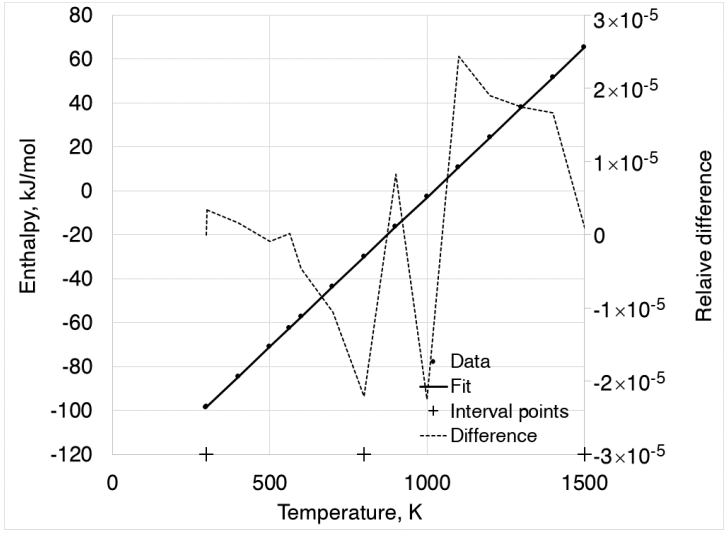


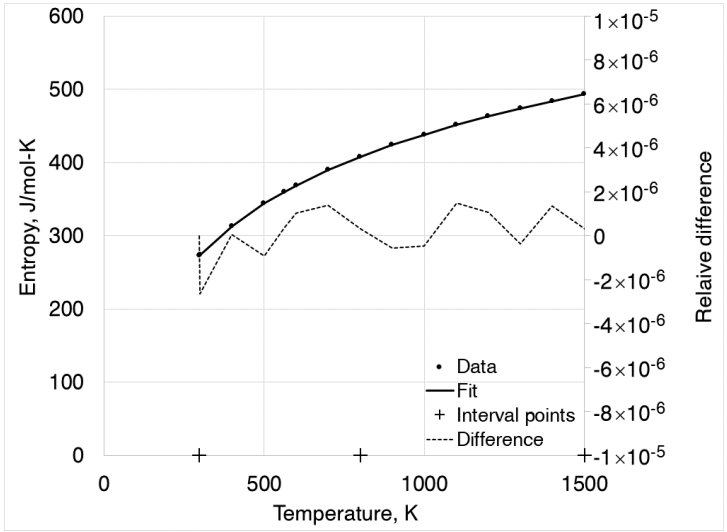


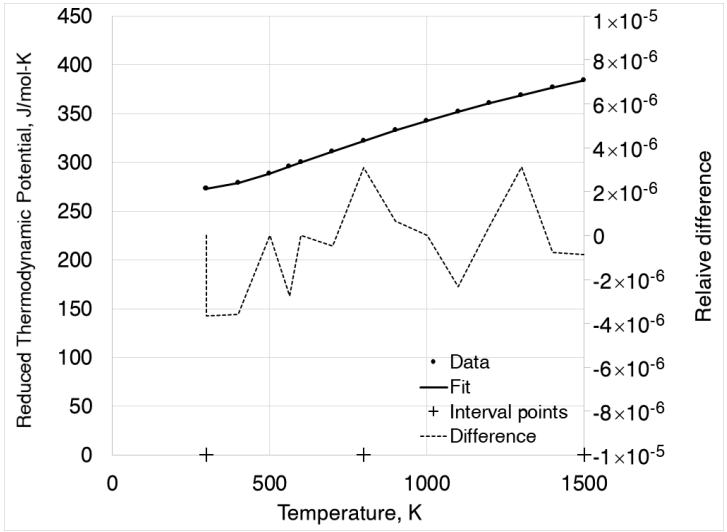
Mercury Iodide (Hg_2I_2) – Hg2I2(l)

MW = 654.99 g/mol – $\Delta_f H^o$(298.15 K) = -98.564 kJ/mol – S^o(298.15 K) = 273.258 J/mol-K – p^o = 0.1 MPa



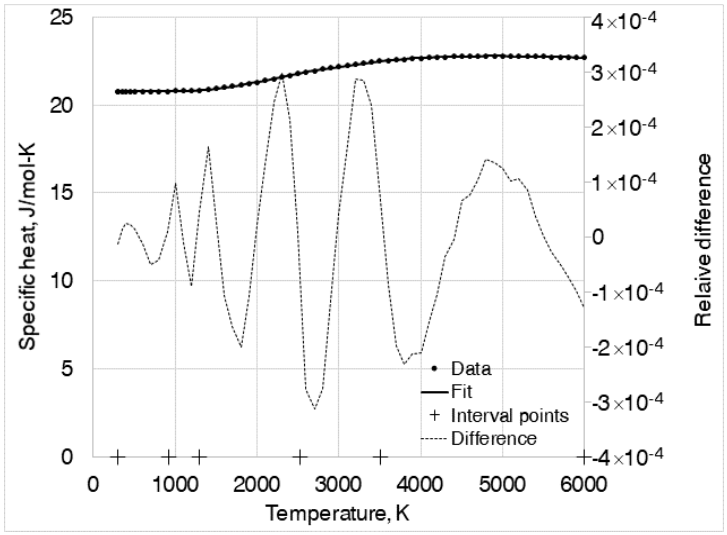


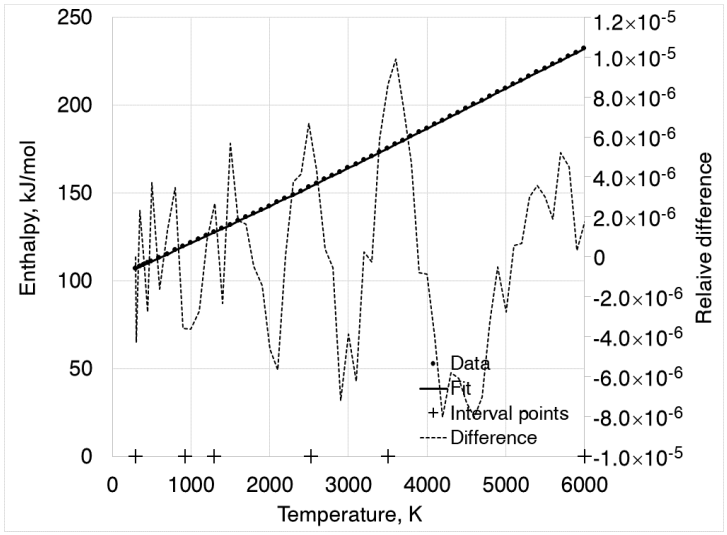


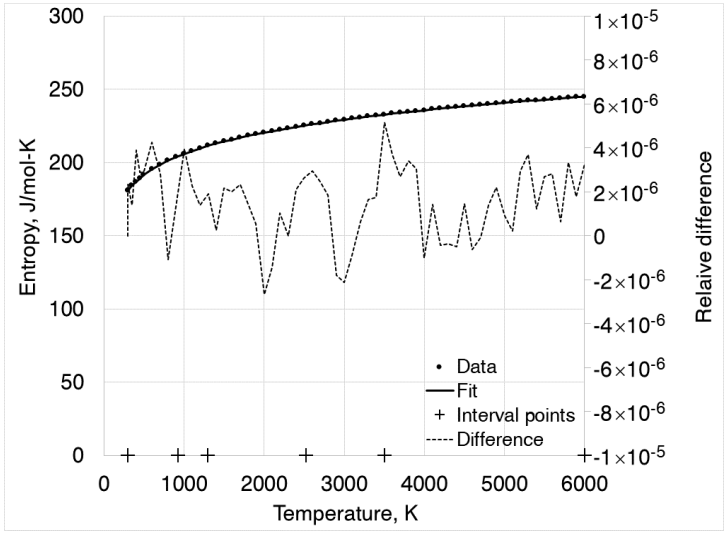


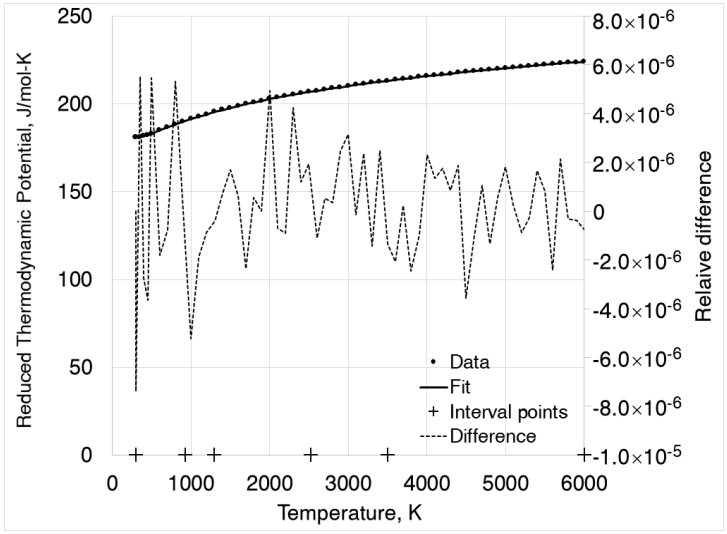
Iodine (I) – I1(g)

MW = 126.9045 g/mol – $\Delta_f H^o$(298.15 K) = 106.762 kJ/mol – S^o(298.15 K) = 180.786 J/mol-K – p^o = 0.1 MPa



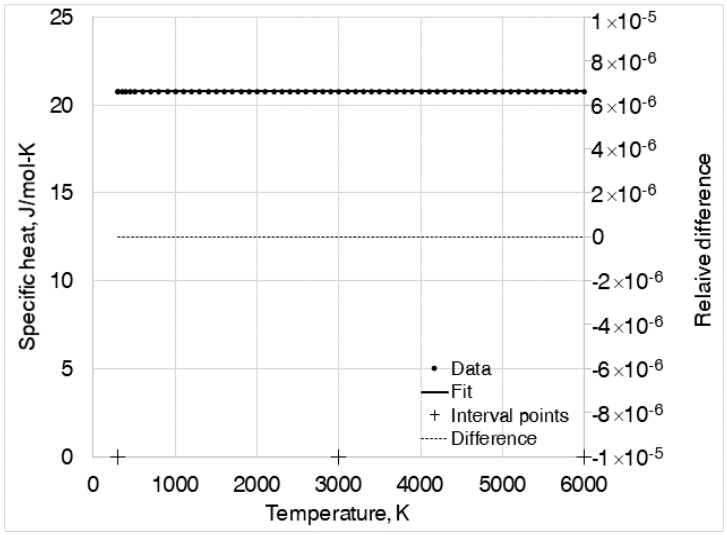


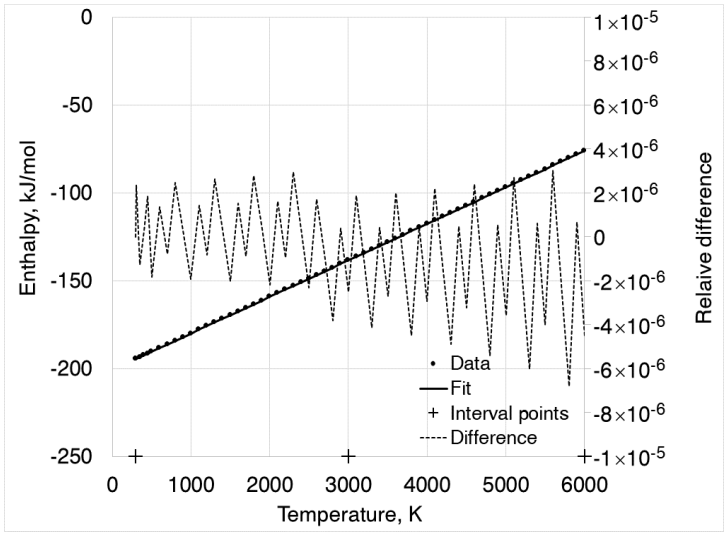


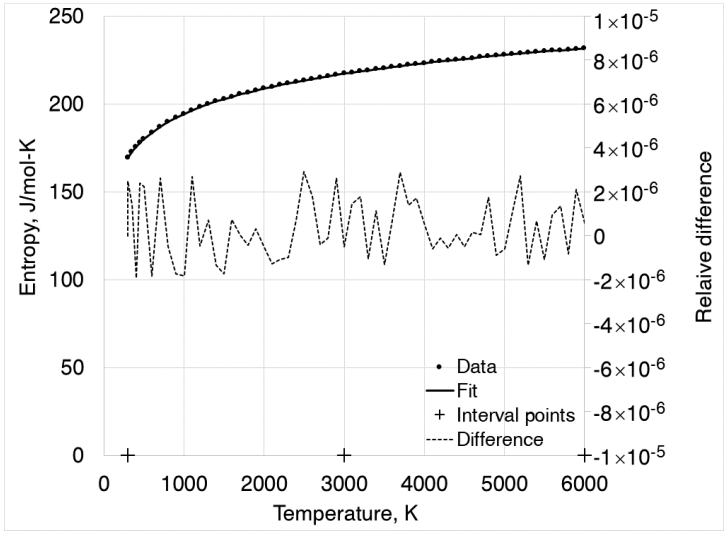


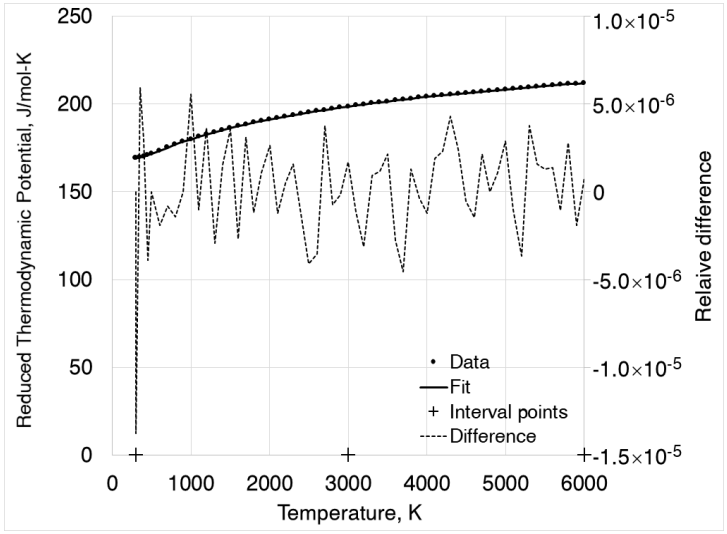
Iodine, Ion (I^-) – I1-(g)

MW = 126.905 g/mol – $\Delta_f H^o$(298.15 K) = -194.591 kJ/mol – S^o(298.15 K) = 169.26 J/mol-K – p^o = 0.1 MPa



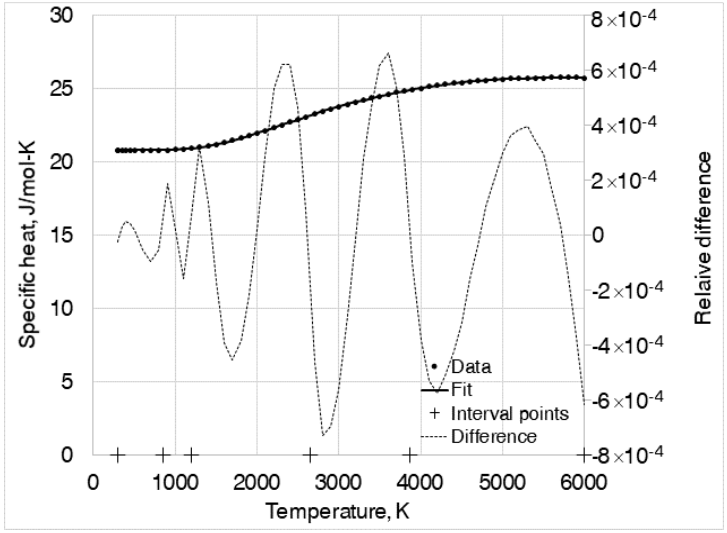


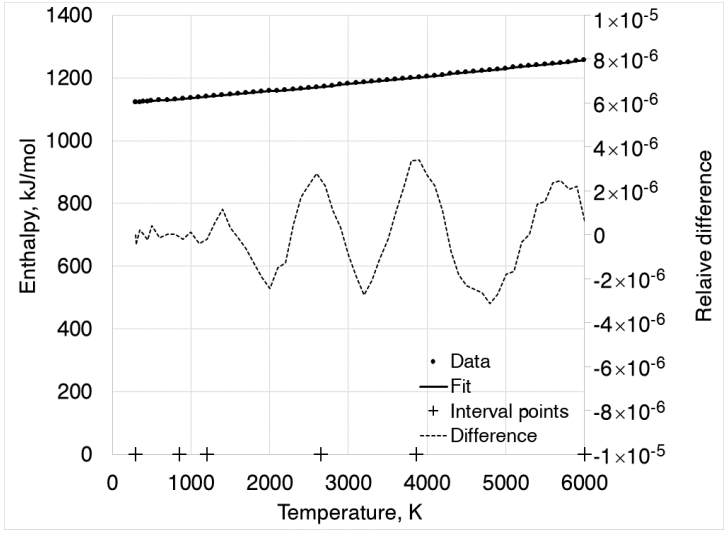


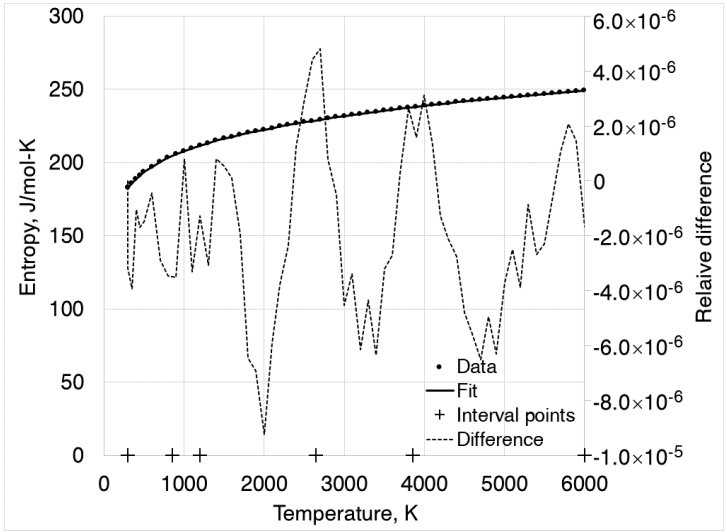


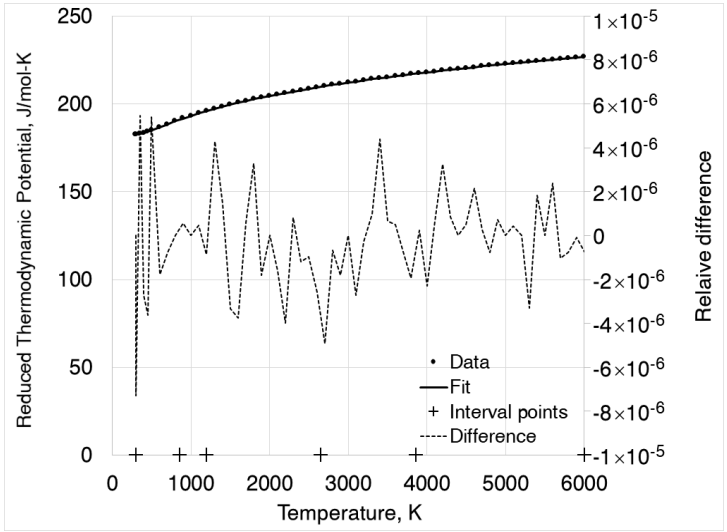
Iodine, Ion (I^+) – I1+(g)

MW = 126.904 g/mol – $\Delta_f H^o$(298.15 K) = 1121.353 kJ/mol – S^o(298.15 K) = 182.642 J/mol-K – p^o = 0.1 MPa



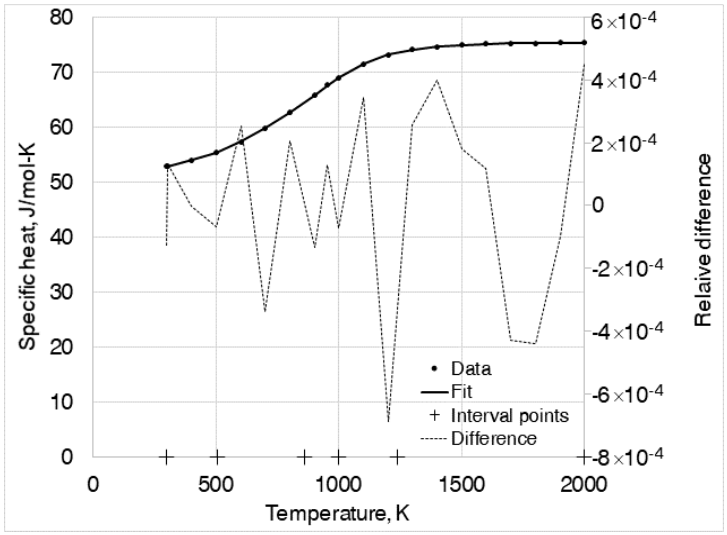


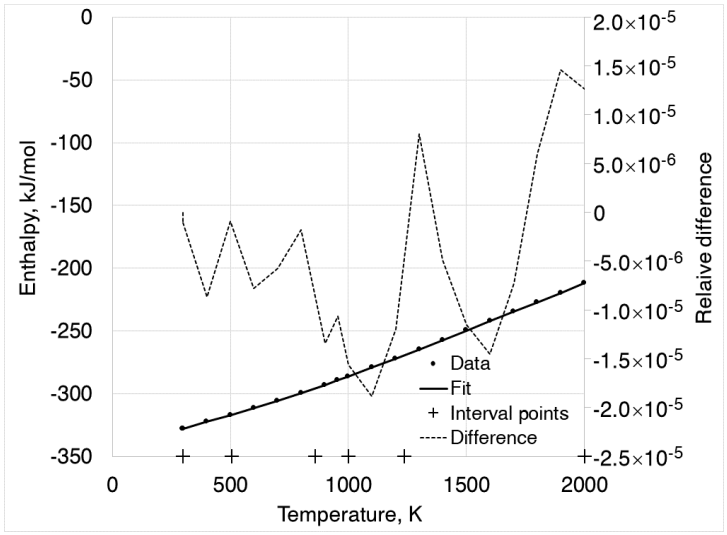


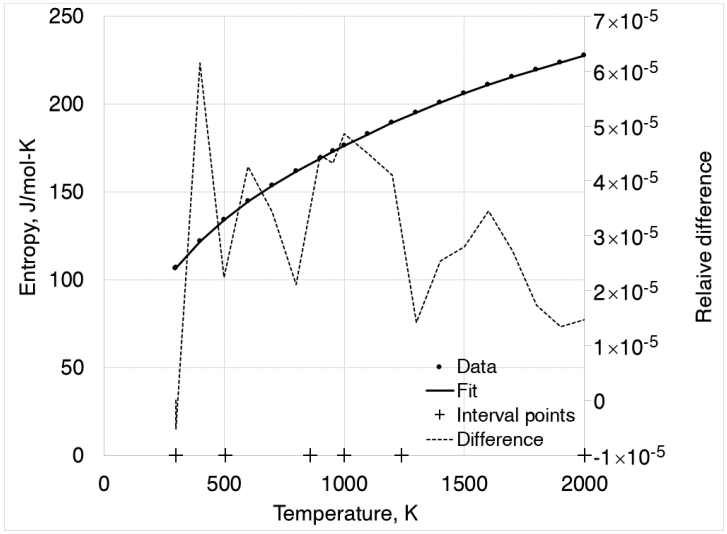


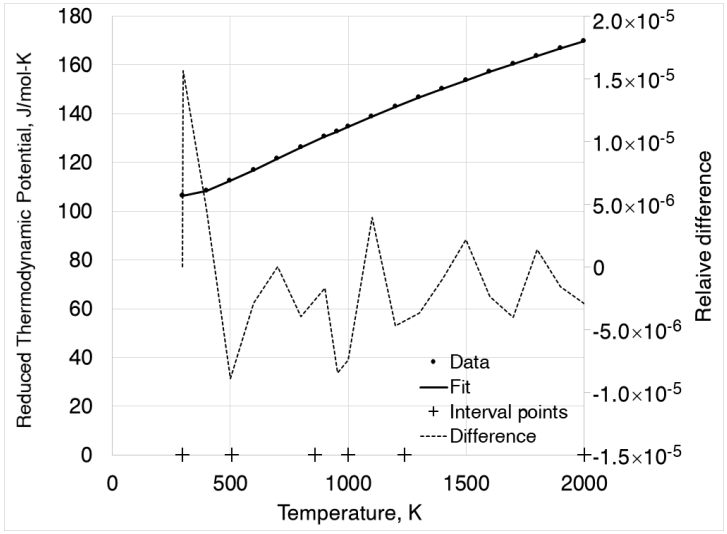
Potassium Iodide (KI) – I1K1(cr)

MW = 166.0028 g/mol – $\Delta_f H^o$(298.15 K) = -327.9 kJ/mol – S^o(298.15 K) = 106.387 J/mol-K – p^o = 0.1 MPa



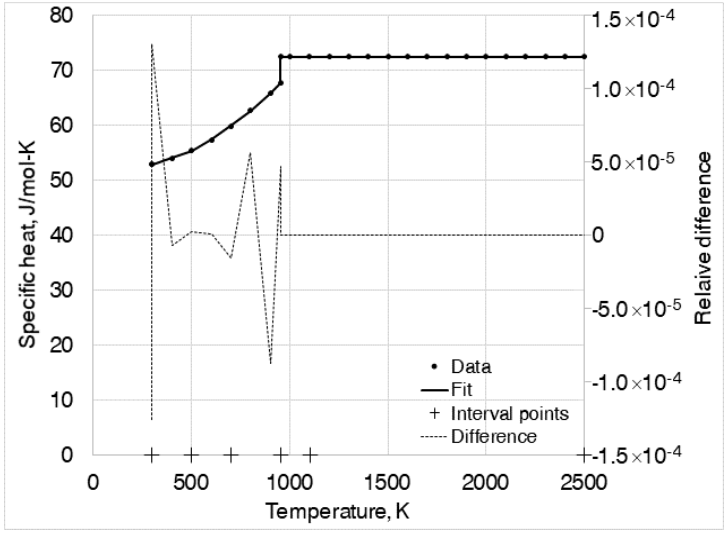


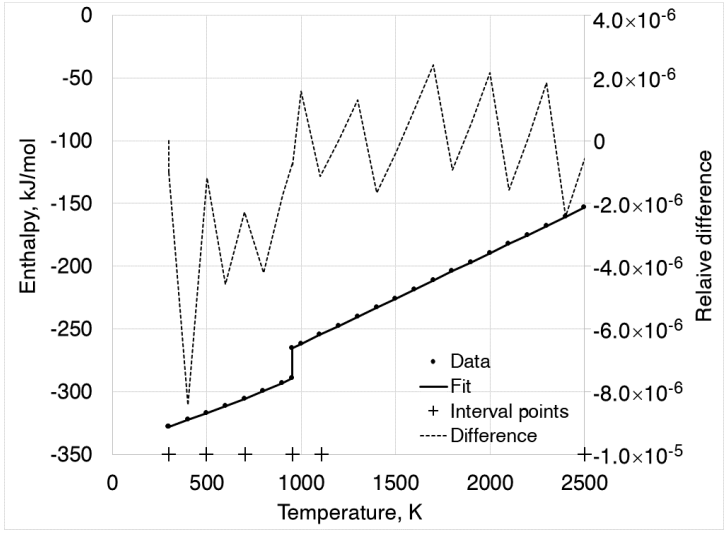


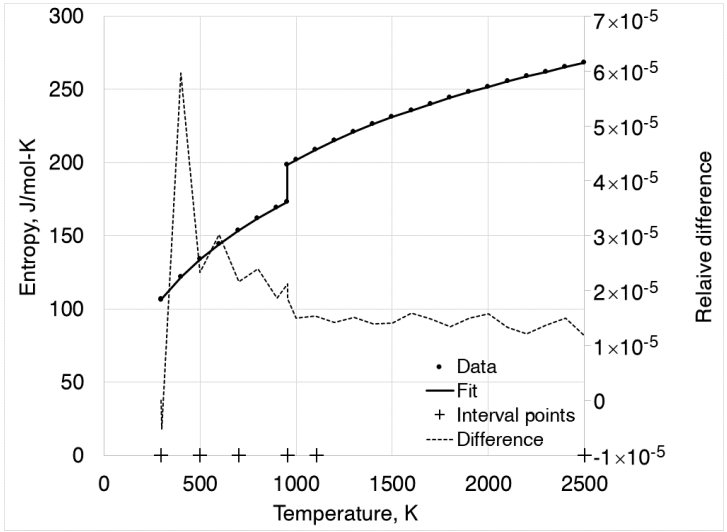


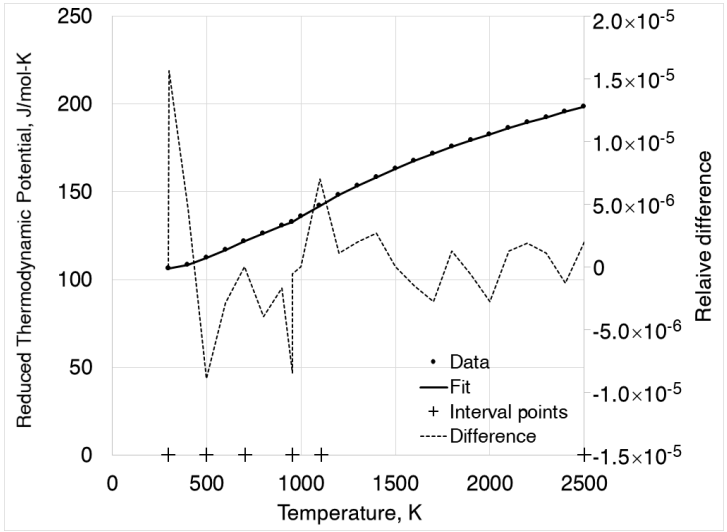
Potassium Iodide (KI) – I1K1(cr,l)

MW = 166.0028 g/mol – $\Delta_f H^o$(298.15 K) = -327.9 kJ/mol – S^o(298.15 K) = 106.387 J/mol-K – p^o = 0.1 MPa



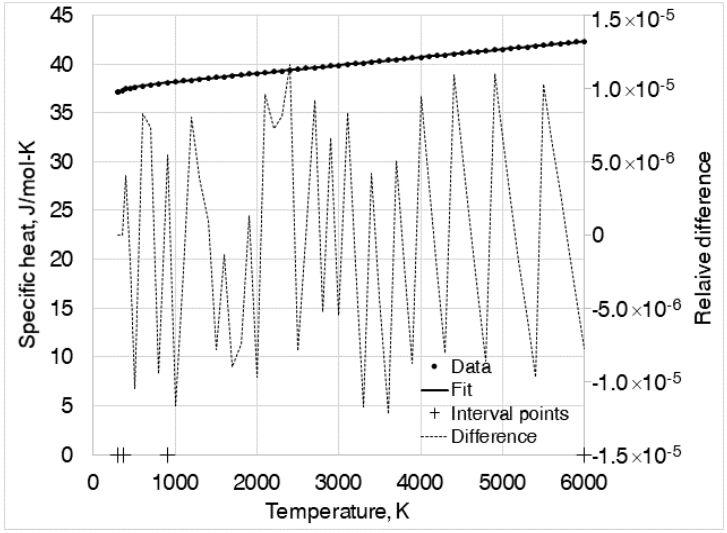


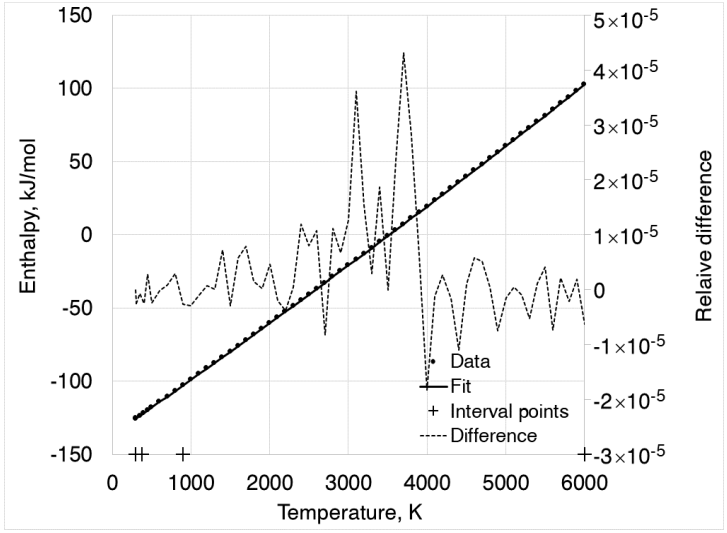


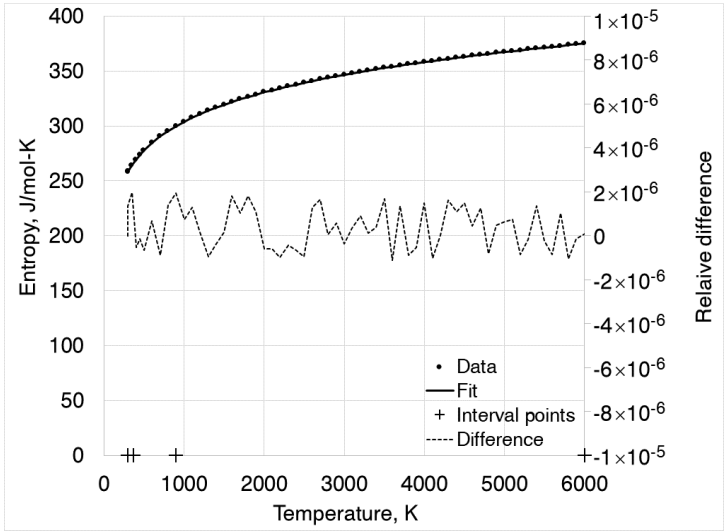


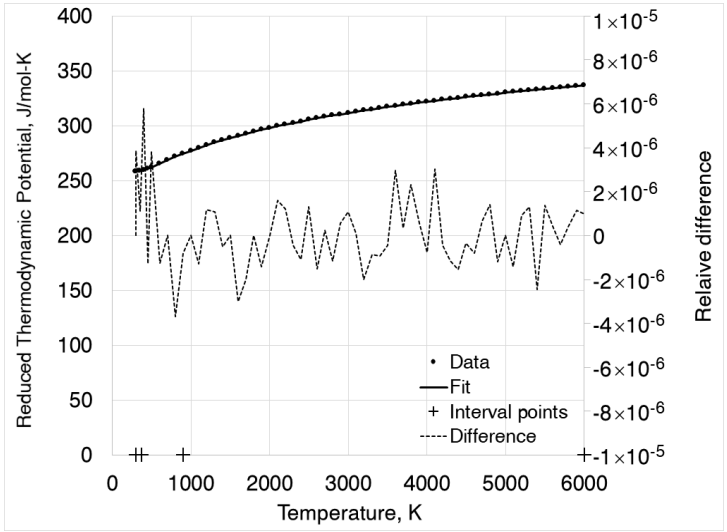
Potassium Iodide (KI) – I1K1(g)

MW = 166.0028 g/mol – $\Delta_f H^o$(298.15 K) = -125.52 kJ/mol – S^o(298.15 K) = 258.283 J/mol-K – p^o = 0.1 MPa



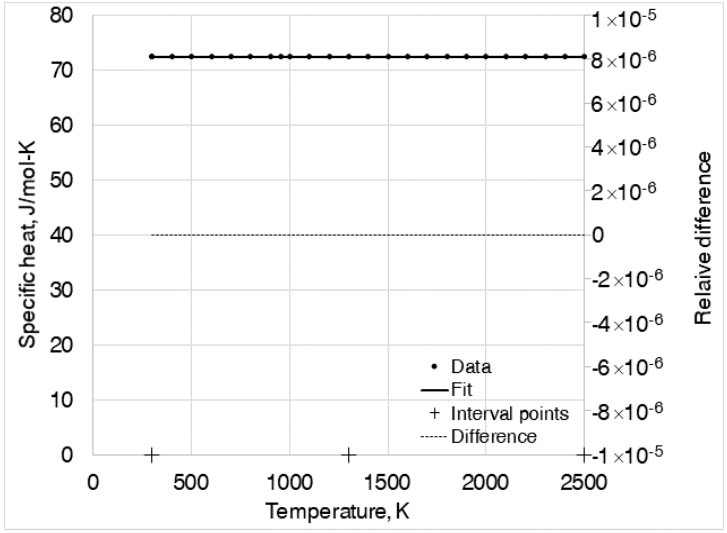


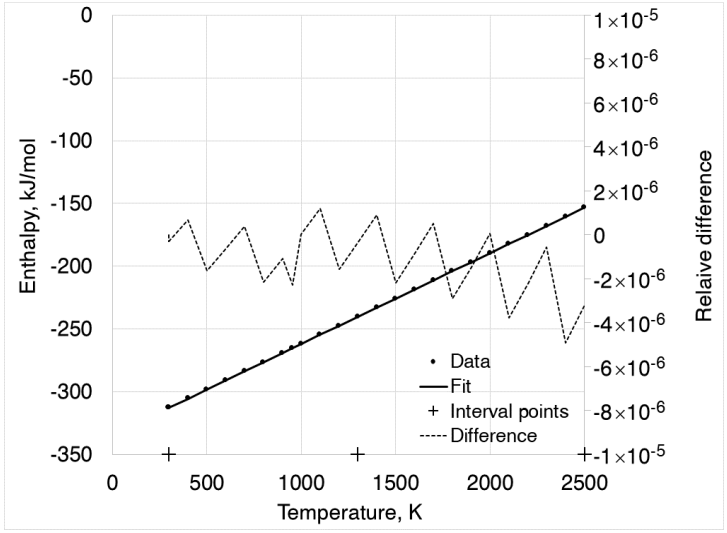


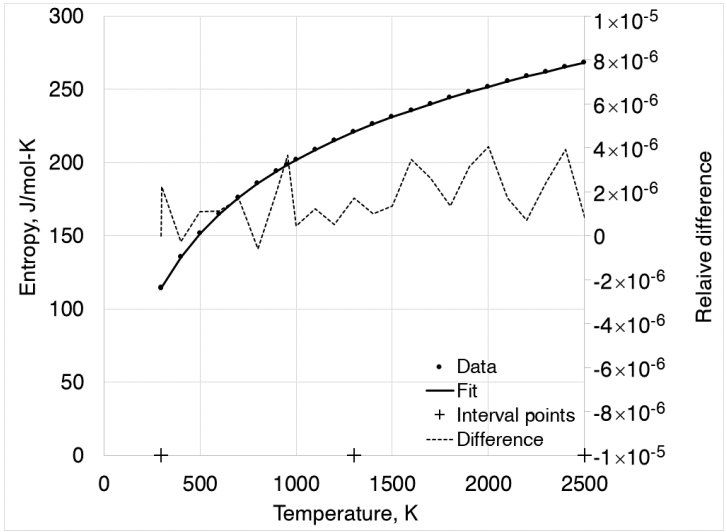


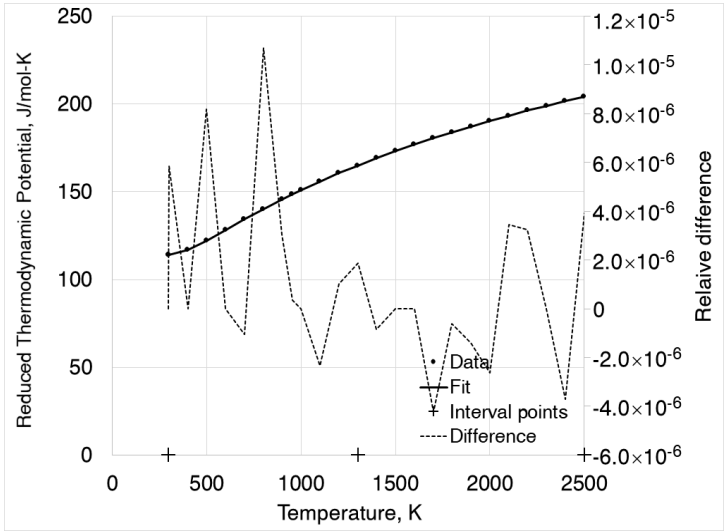
Potassium Iodide (KI) – I1K1(l)

MW = 166.0028 g/mol – $\Delta_f H^o$(298.15 K) = -312.846 kJ/mol – S^o(298.15 K) = 114.098 J/mol-K – p^o = 0.1 MPa



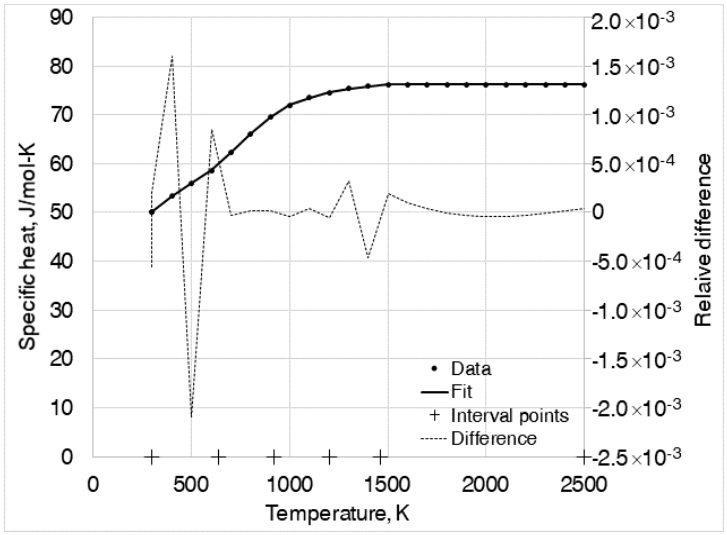


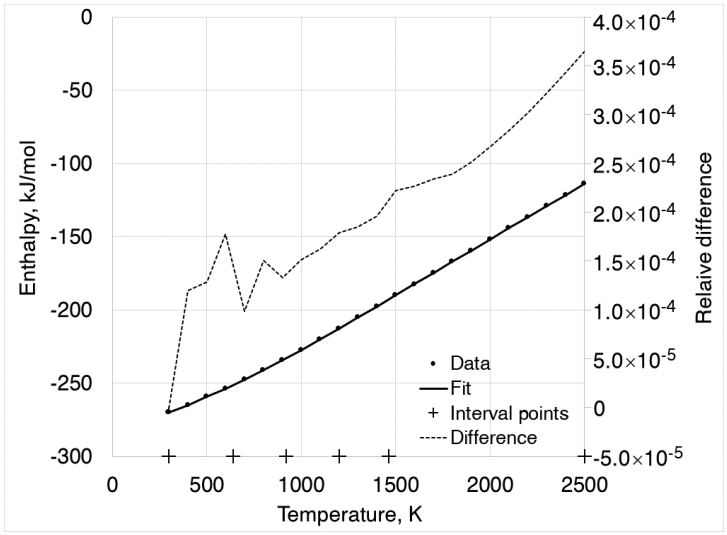


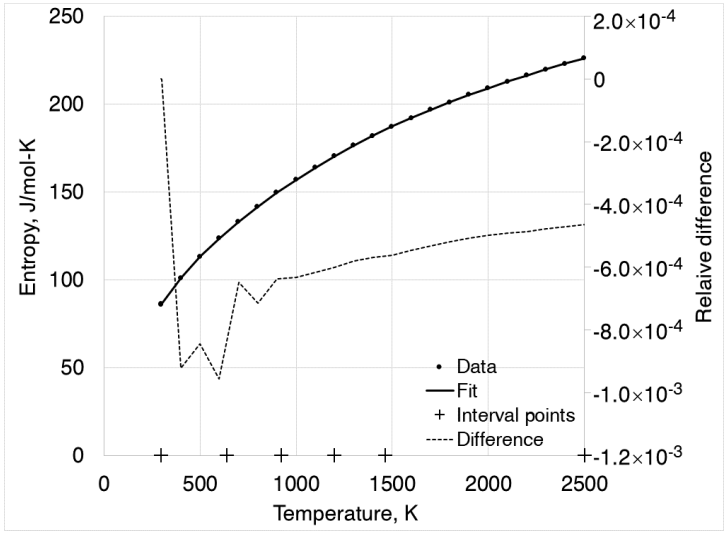


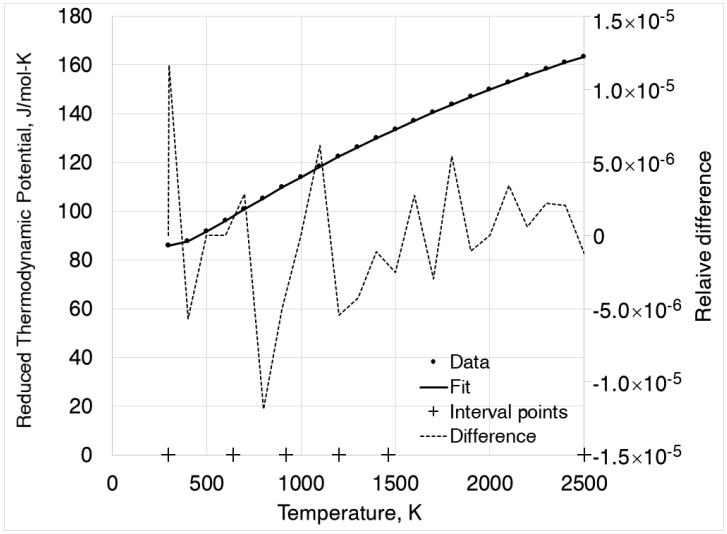
Lithium Iodide (LiI) – I1Li1(cr)

MW = 133.846 g/mol – $\Delta_f H^o$(298.15 K) = -270.077 kJ/mol – S^o(298.15 K) = 85.772 J/mol-K – p^o = 0.1 MPa



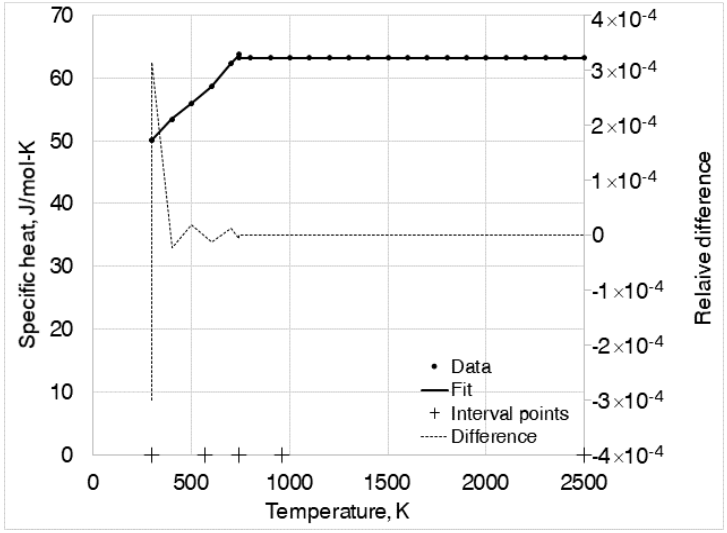


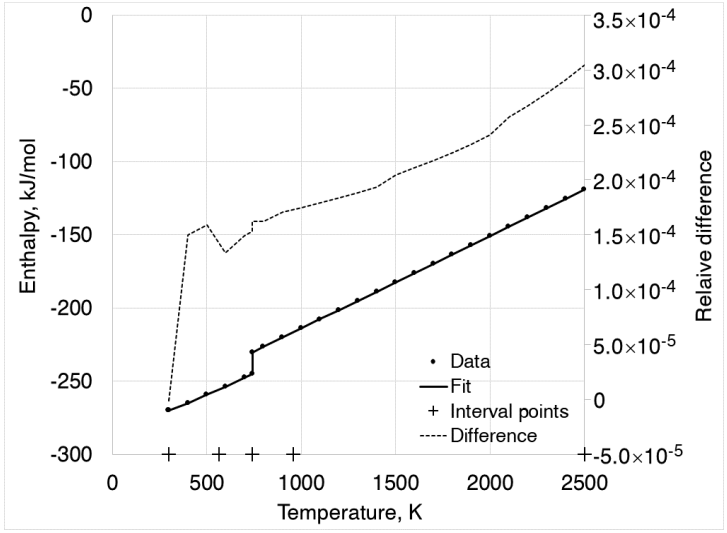


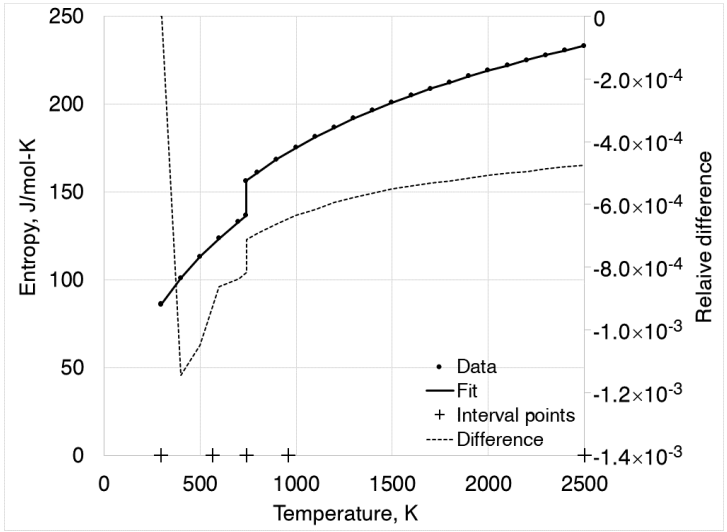


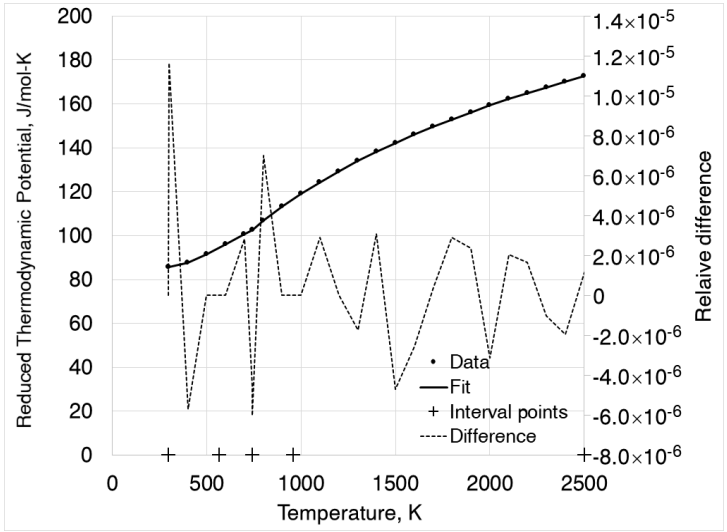
Lithium Iodide (LiI) – I1Li1(cr,l)

MW = 133.846 g/mol – $\Delta_f H^o$(298.15 K) = -270.077 kJ/mol – S^o(298.15 K) = 85.772 J/mol-K – p^o = 0.1 MPa



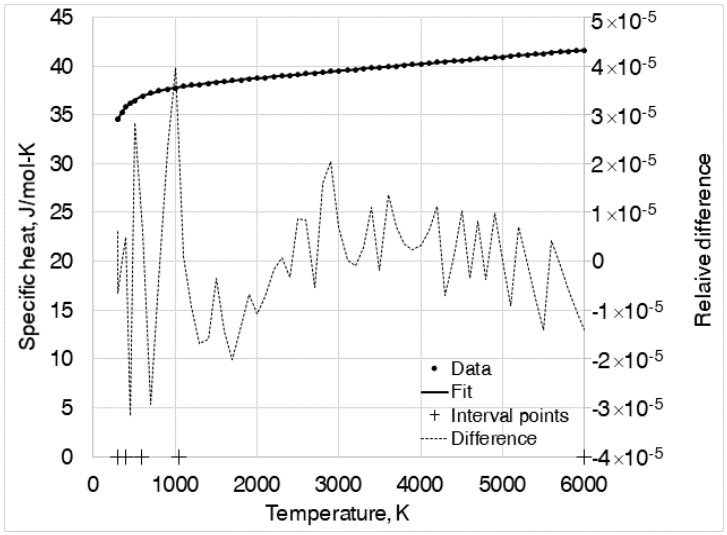


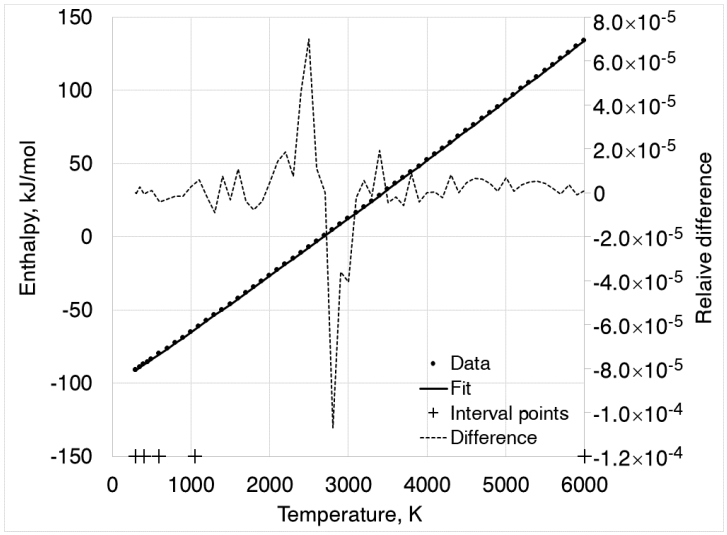


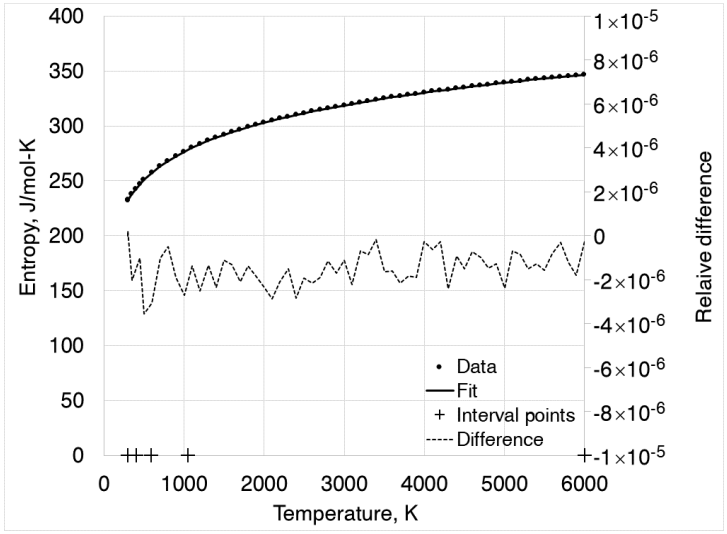


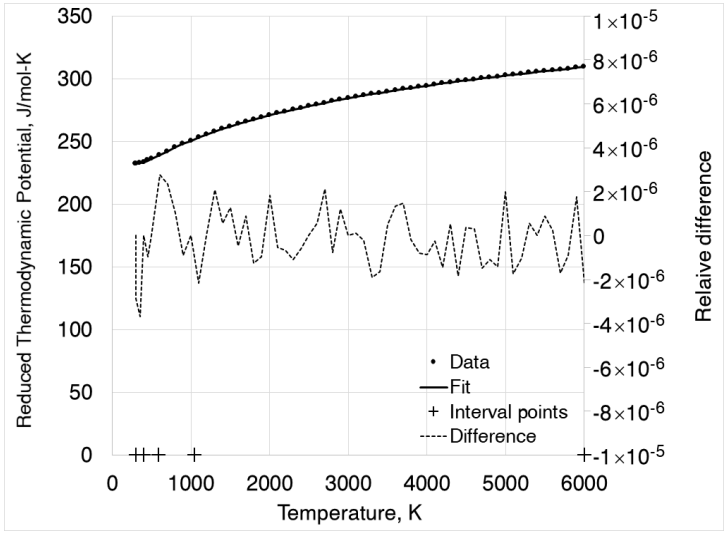
Lithium Iodide (LiI) – I1Li1(g)

MW = 133.846 g/mol – $\Delta_f H^o$(298.15 K) = -91.002 kJ/mol – S^o(298.15 K) = 232.262 J/mol-K – p^o = 0.1 MPa



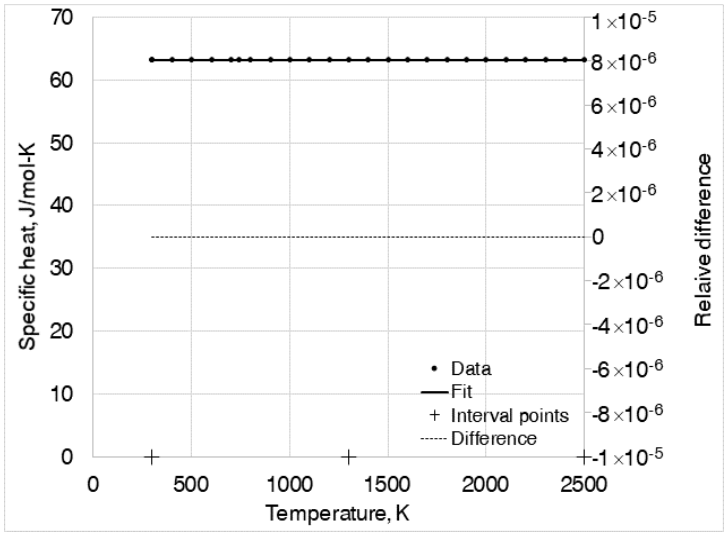


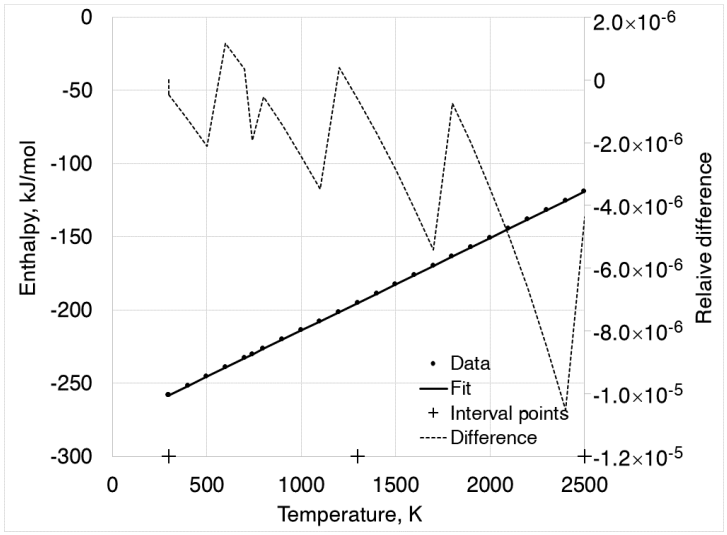


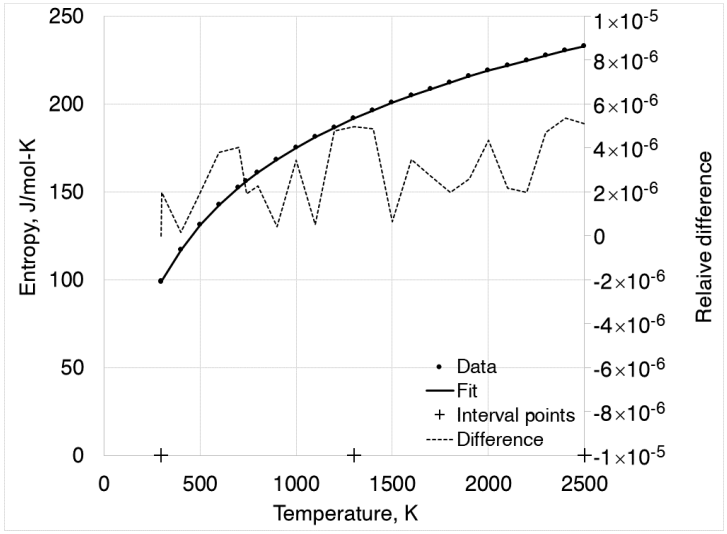


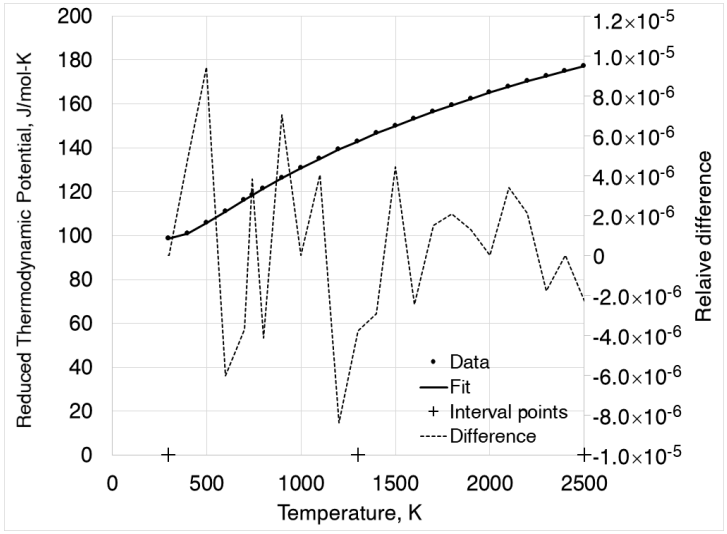
Lithium Iodide (LiI) – I1Li1(l)

MW = 133.846 g/mol – $\Delta_f H^o$(298.15 K) = -258.349 kJ/mol – S^o(298.15 K) = 98.611 J/mol-K – p^o = 0.1 MPa



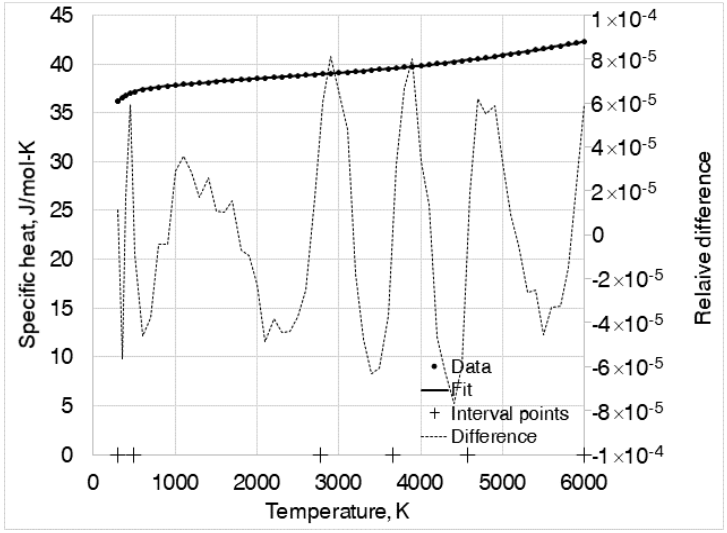


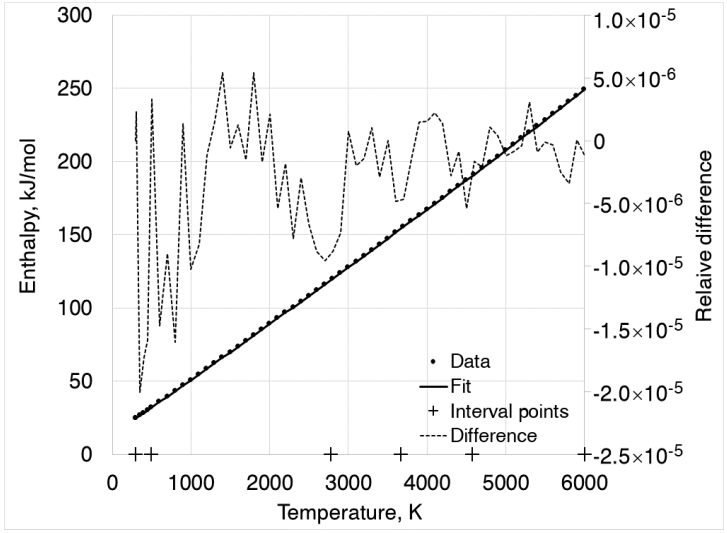


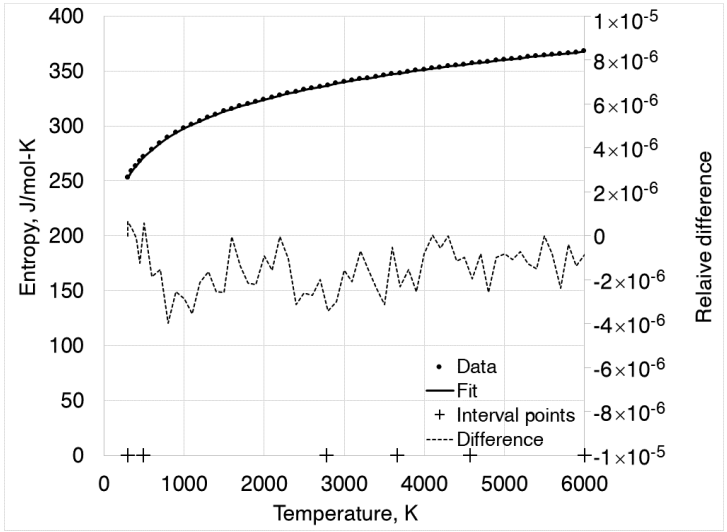


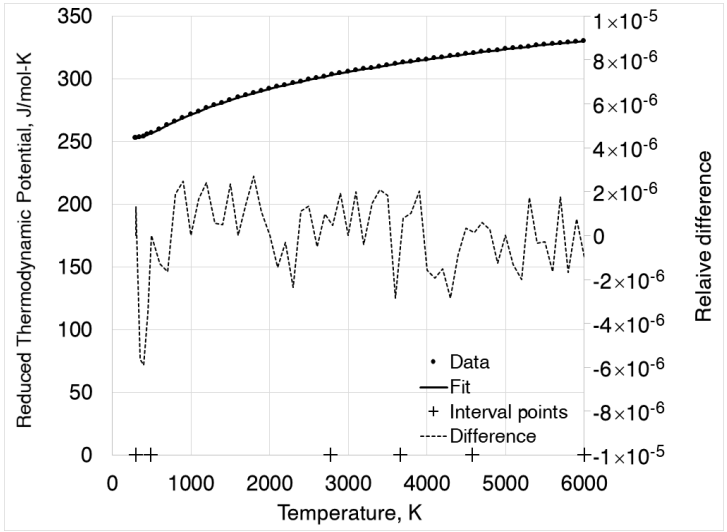
Magnesium Iodide (MgI) – I1Mg1(g)

MW = 151.21 g/mol – $\Delta_f H^o$(298.15 K) = 24.614 kJ/mol – S^o(298.15 K) = 252.647 J/mol-K – p^o = 0.1 MPa



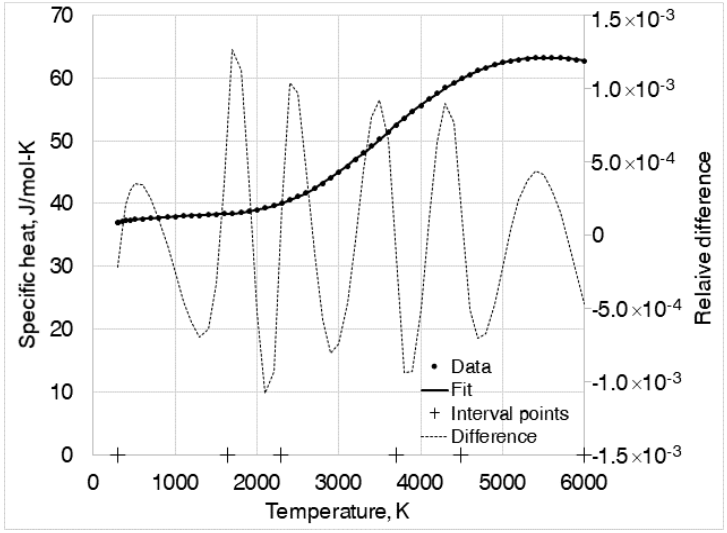


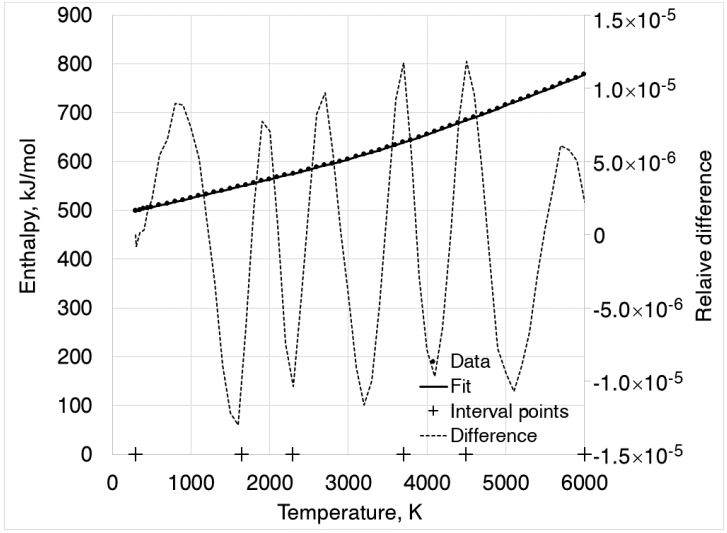


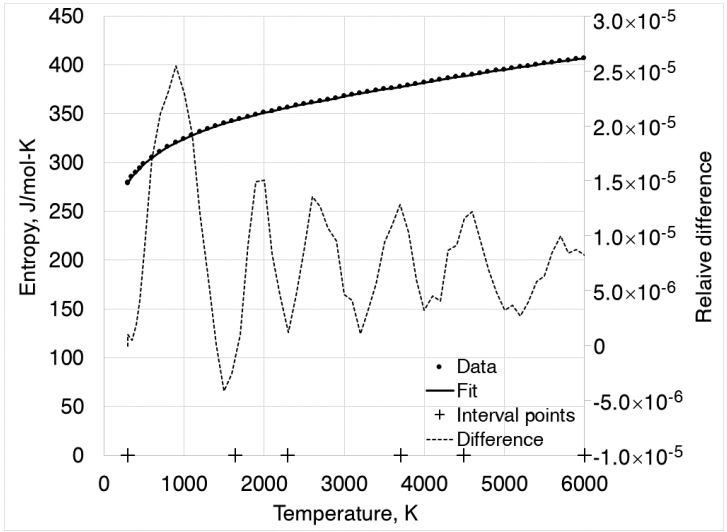


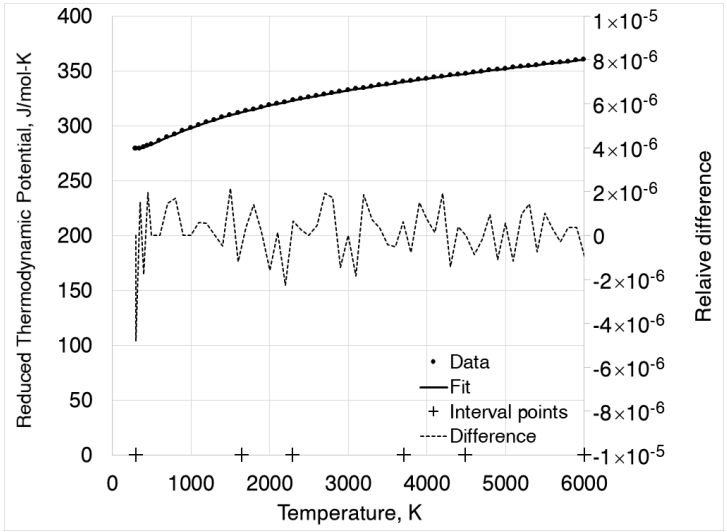
Molybdenum Iodide (MoI) – I1Mo1(g)

MW = 222.84 g/mol – $\Delta_f H^o$(298.15 K) = 498.733 kJ/mol – S^o(298.15 K) = 278.974 J/mol-K – p^o = 0.1 MPa



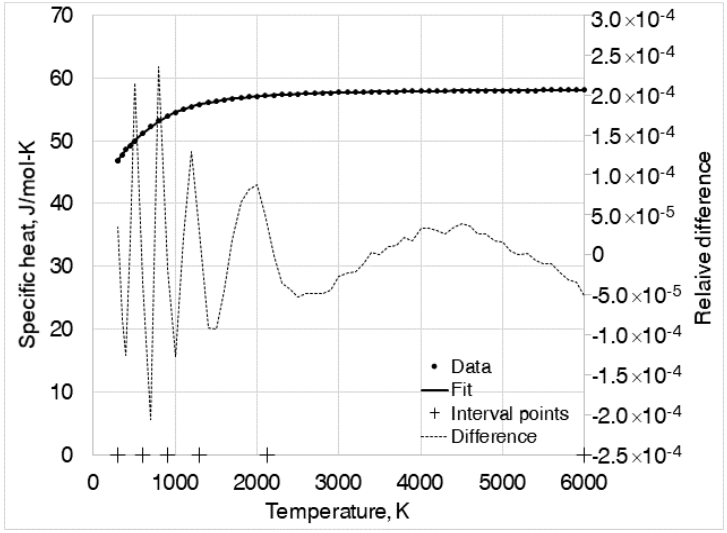


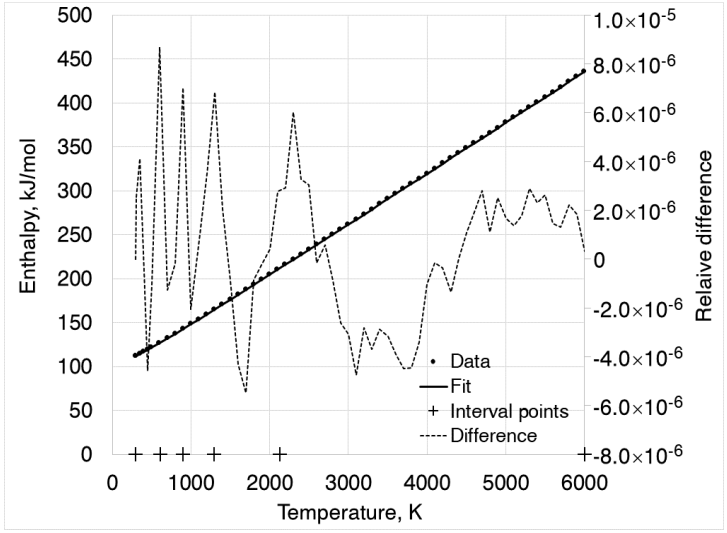


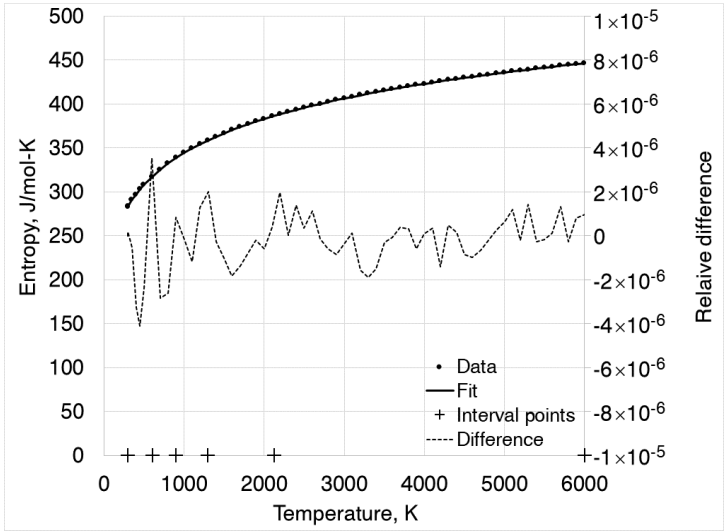


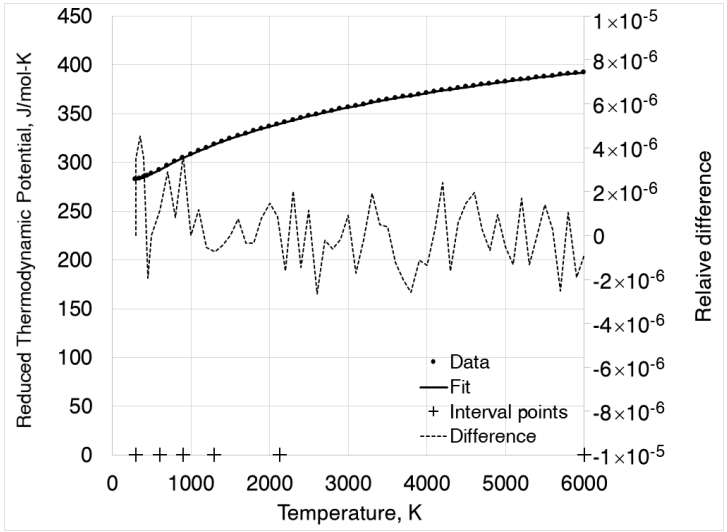
Nitrosyl Iodide (ONI) – I1N1O1(g)

MW = 156.9106 g/mol – $\Delta_f H^o$(298.15 K) = 112.131 kJ/mol – S^o(298.15 K) = 282.96 J/mol-K – p^o = 0.1 MPa



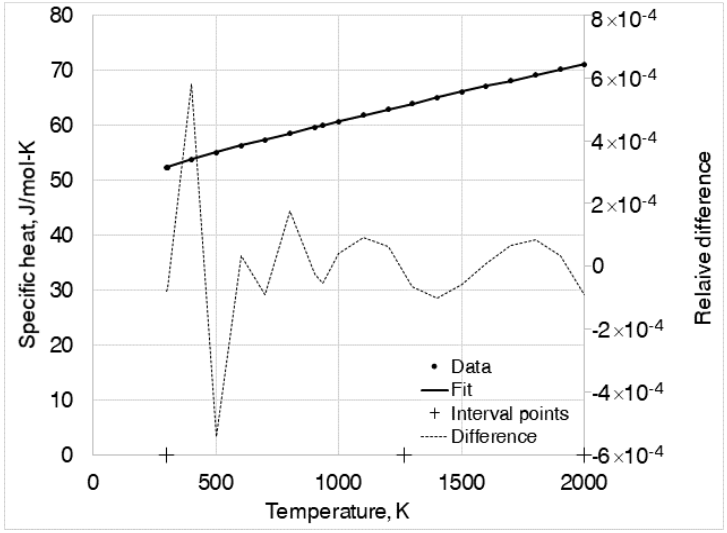


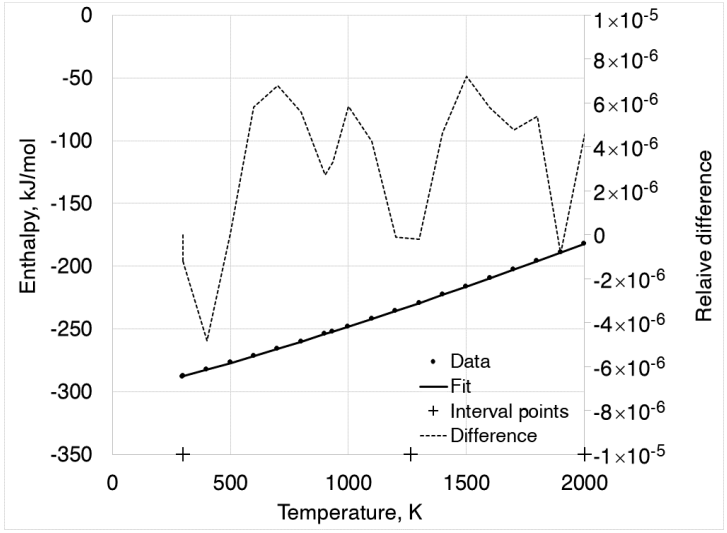


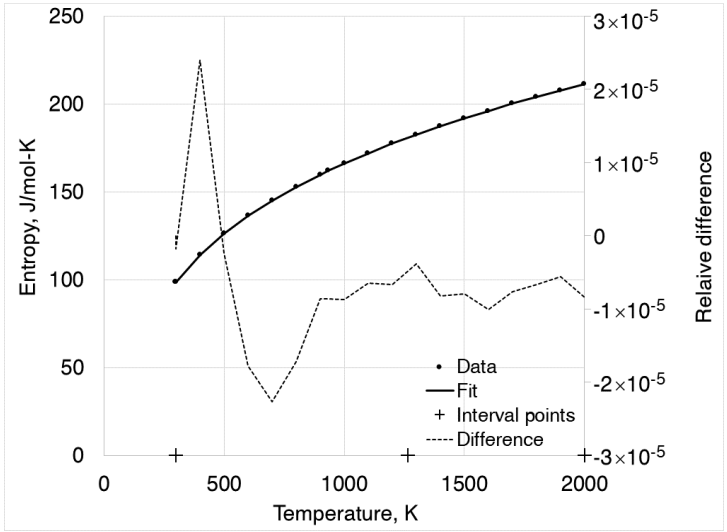


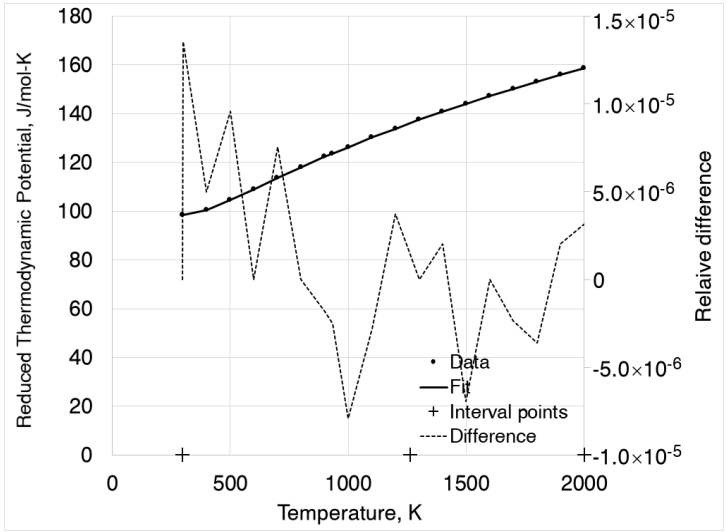
Sodium Iodide (NaI) – I1Na1(cr)

MW = 149.8943 g/mol – $\Delta_f H^o$(298.15 K) = -287.859 kJ/mol – S^o(298.15 K) = 98.5 J/mol-K – p^o = 0.1 MPa



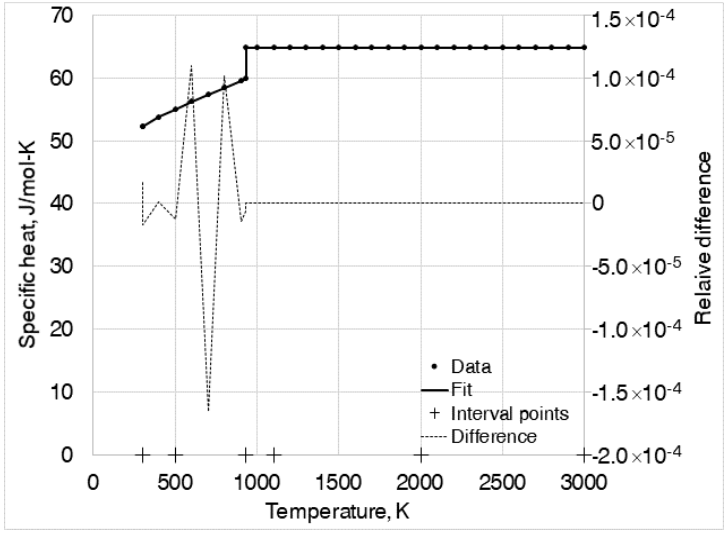


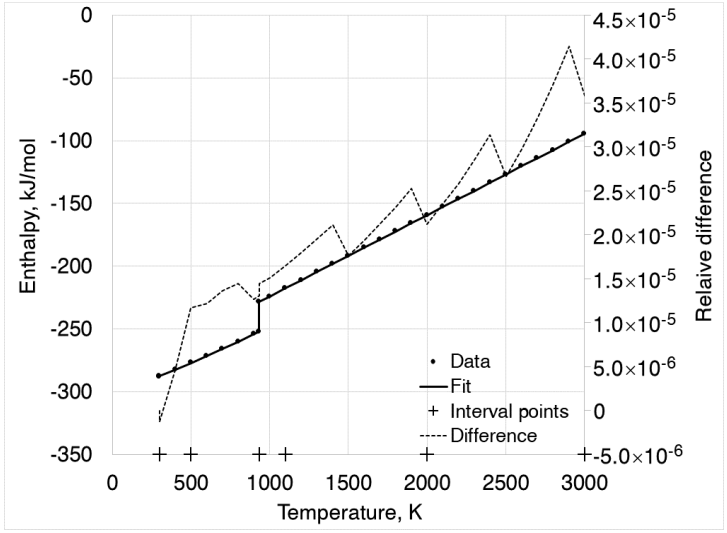


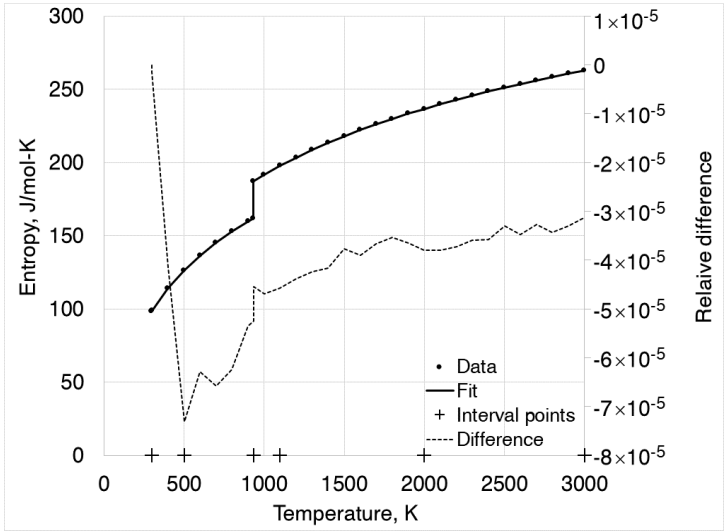


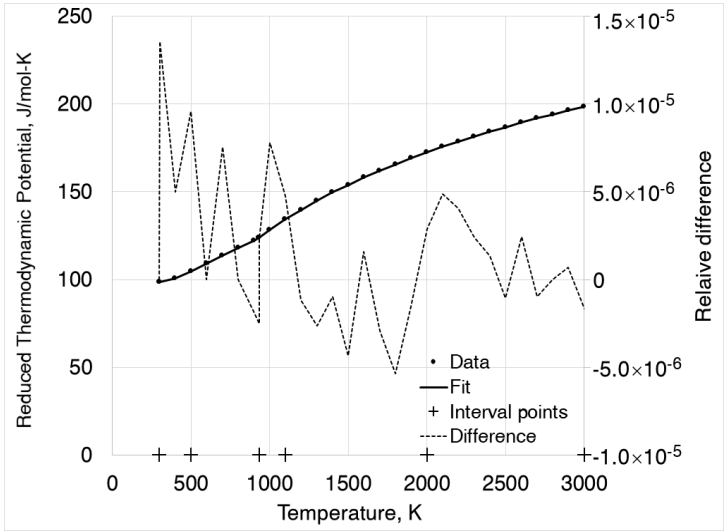
Sodium Iodide (NaI) – I1Na1(cr,l)

MW = 149.8943 g/mol – $\Delta_f H^o$(298.15 K) = -287.859 kJ/mol – S^o(298.15 K) = 98.5 J/mol-K – p^o = 0.1 MPa



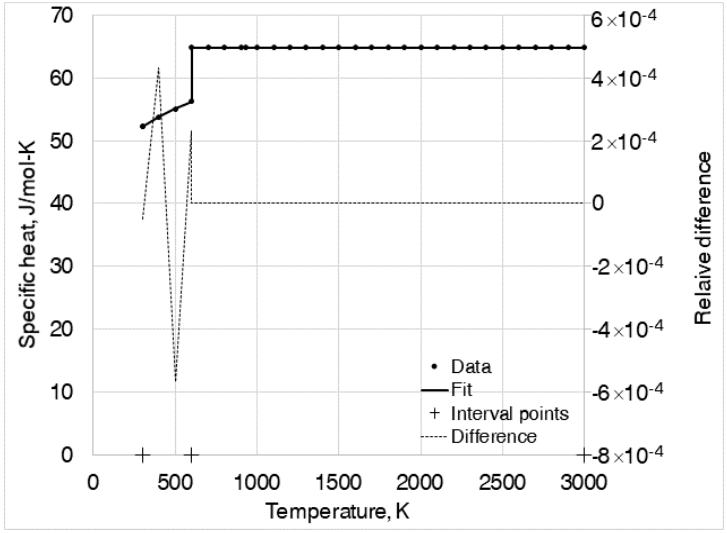


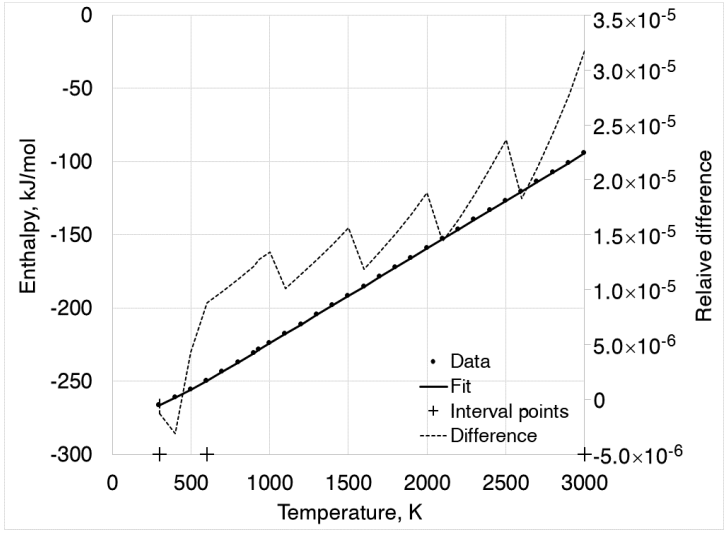


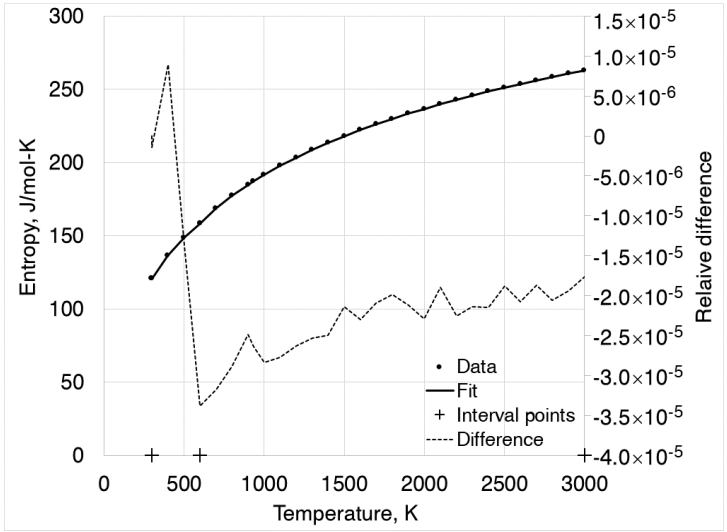


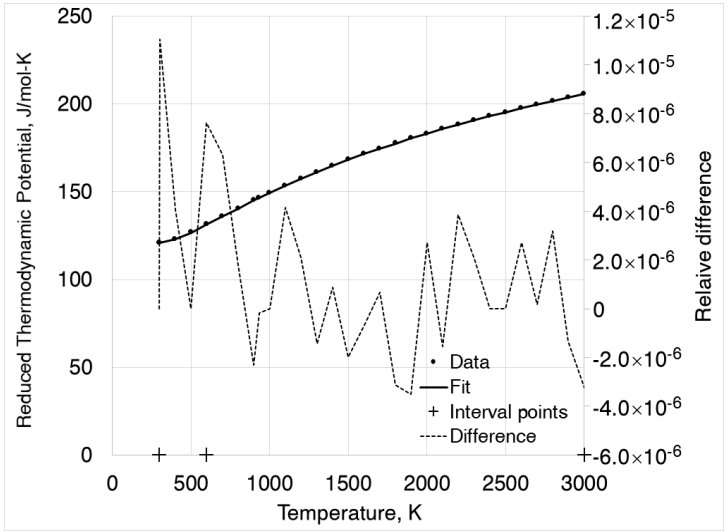
Sodium Iodide (NaI) – I1Na1(l)

MW = 149.8943 g/mol – $\Delta_f H^o$(298.15 K) = -266.508 kJ/mol – S^o(298.15 K) = 120.754 J/mol-K – p^o = 0.1 MPa



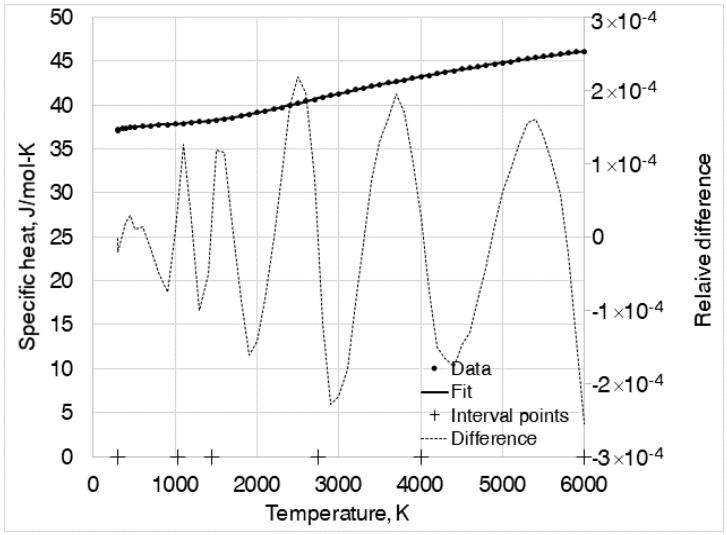


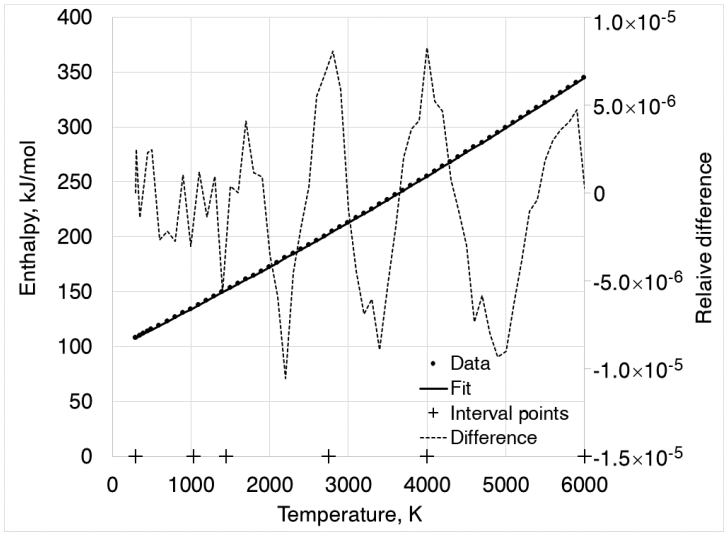


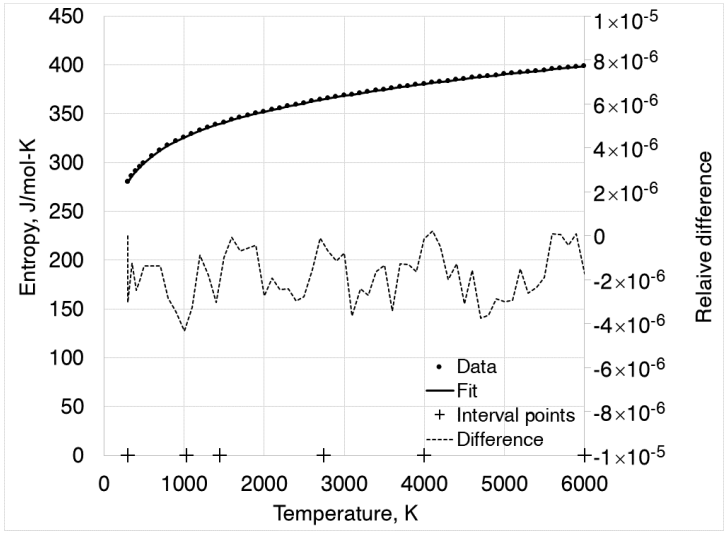


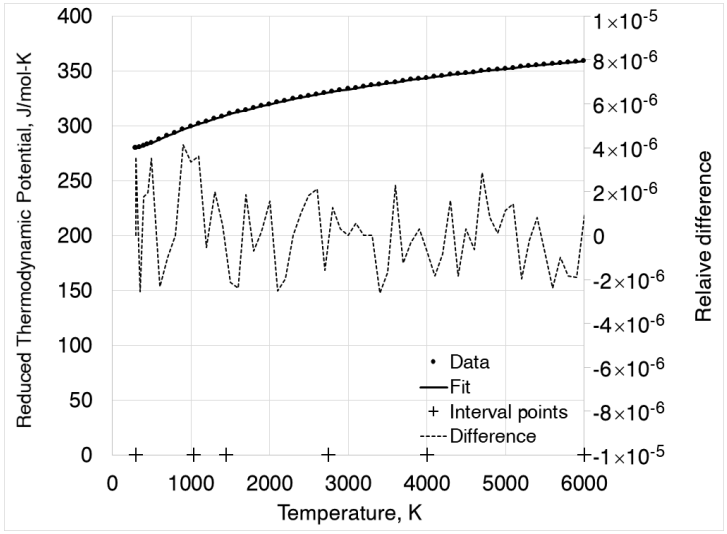
Lead Iodide (PbI) – I1Pb1(g)

MW = 334.1 g/mol – $\Delta_f H^o$(298.15 K) = 107.738 kJ/mol – S^o(298.15 K) = 280.094 J/mol-K – p^o = 0.1 MPa



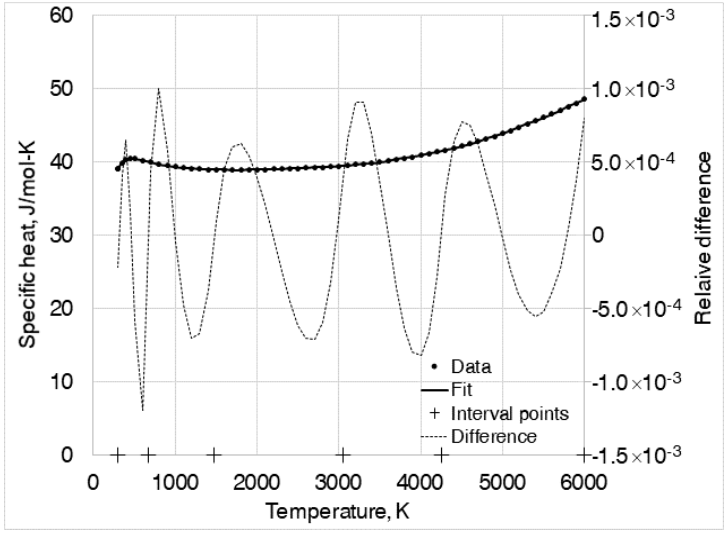


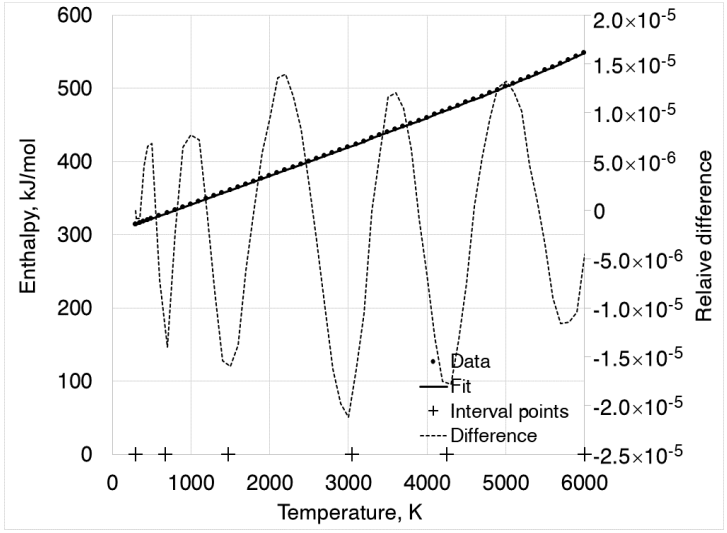


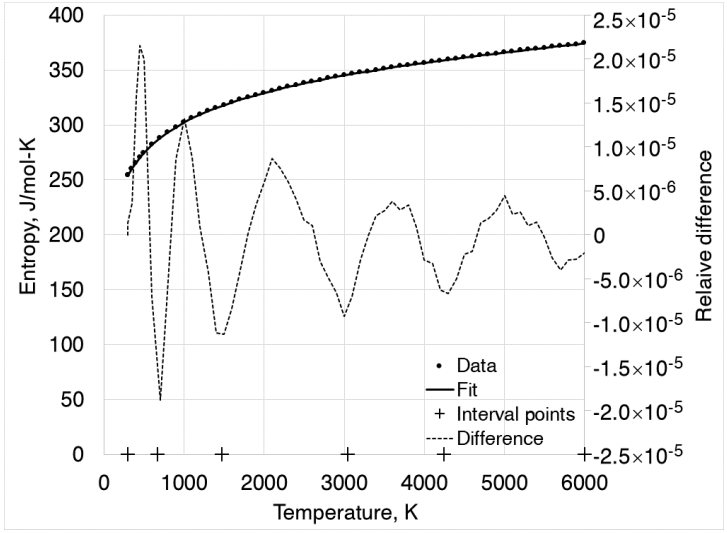


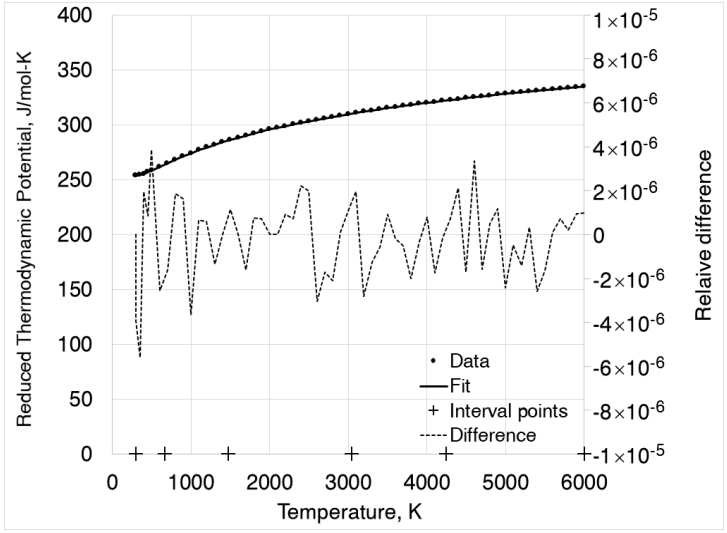
Iodosilylidyne (SiI) – I1Si1(g)

MW = 154.99 g/mol – $\Delta_f H^o$(298.15 K) = 313.629 kJ/mol – S^o(298.15 K) = 253.854 J/mol-K – p^o = 0.1 MPa



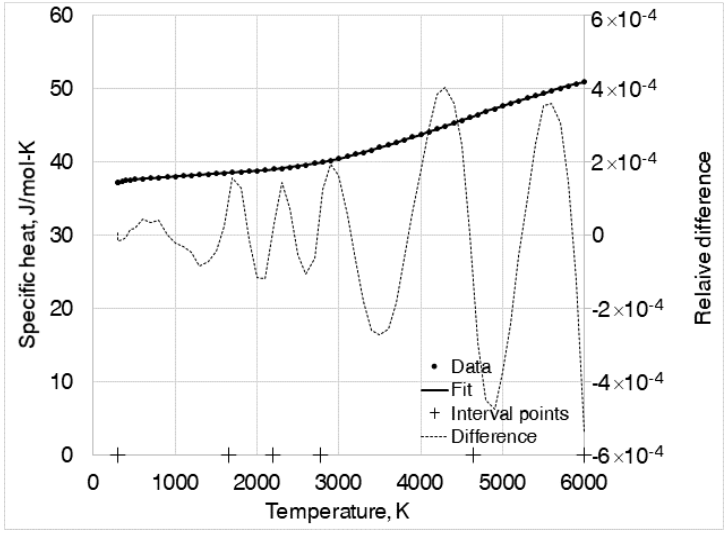


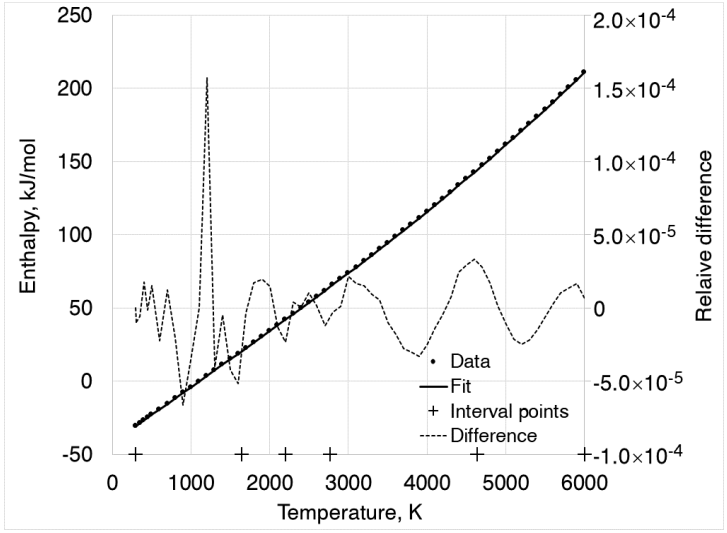


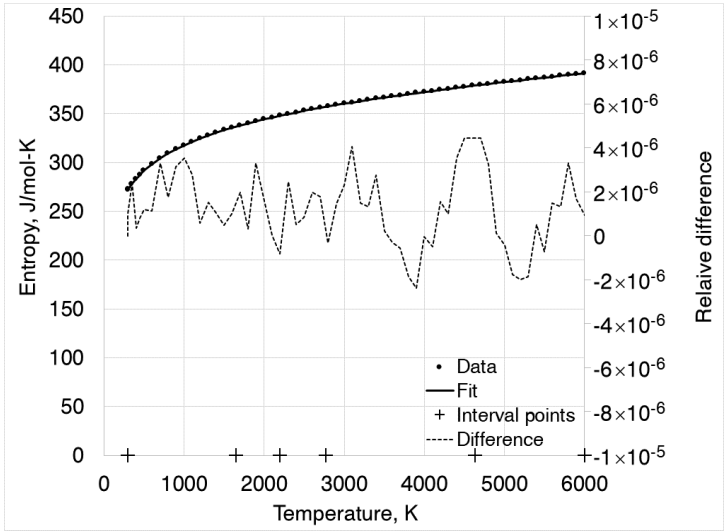


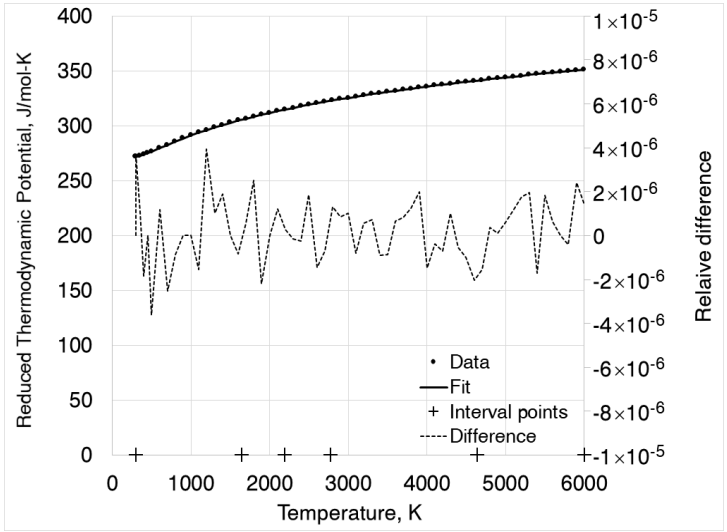
Strontium Iodide (SrI) – I1Sr1(g)

MW = 214.52 g/mol – $\Delta_f H^o$(298.15 K) = -30.494 kJ/mol – S^o(298.15 K) = 272.213 J/mol-K – p^o = 0.1 MPa



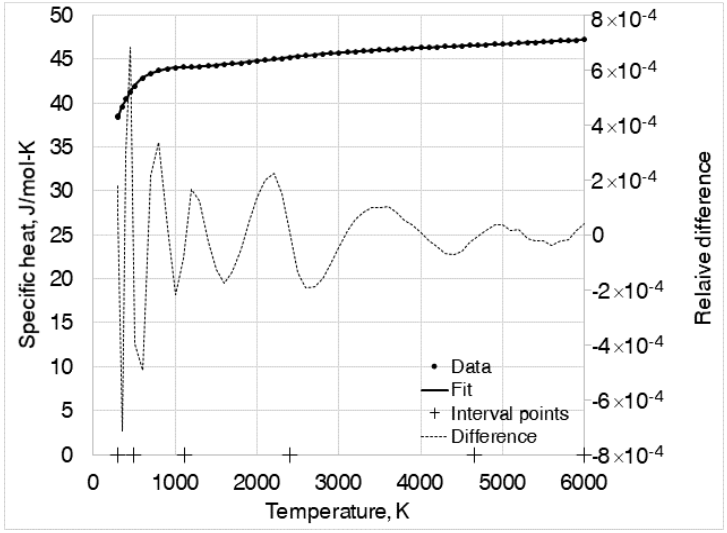


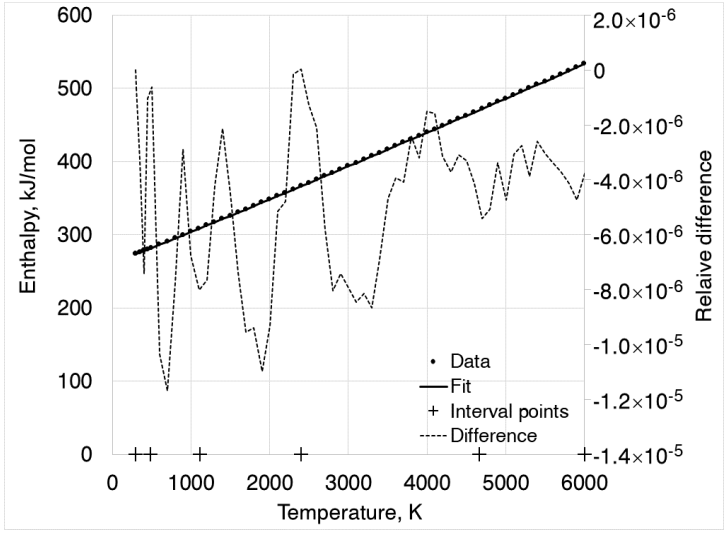


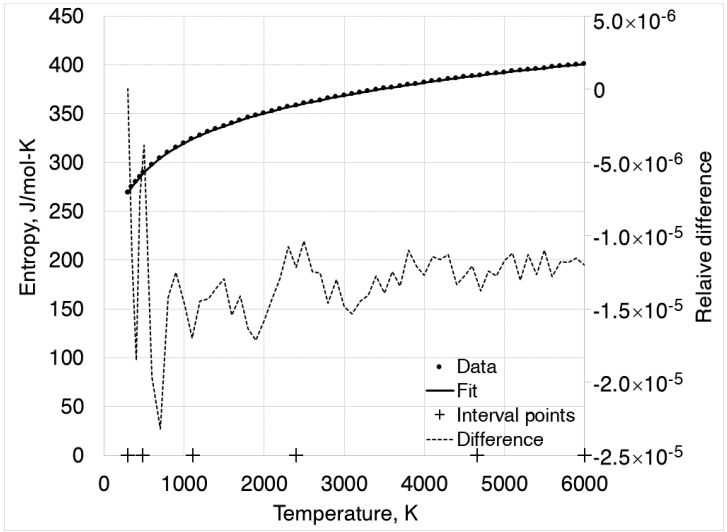


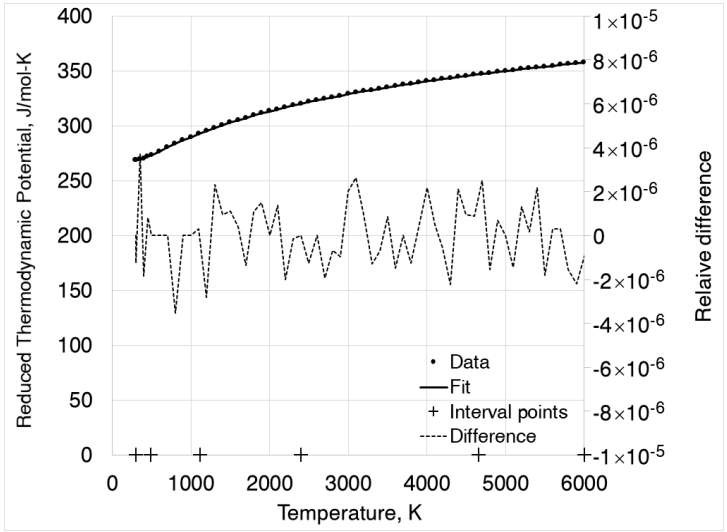
Titanium Iodide (TiI) – I1Ti1(g)

MW = 174.78 g/mol – $\Delta_f H^o$(298.15 K) = 274.052 kJ/mol – S^o(298.15 K) = 268.827 J/mol-K – p^o = 0.1 MPa



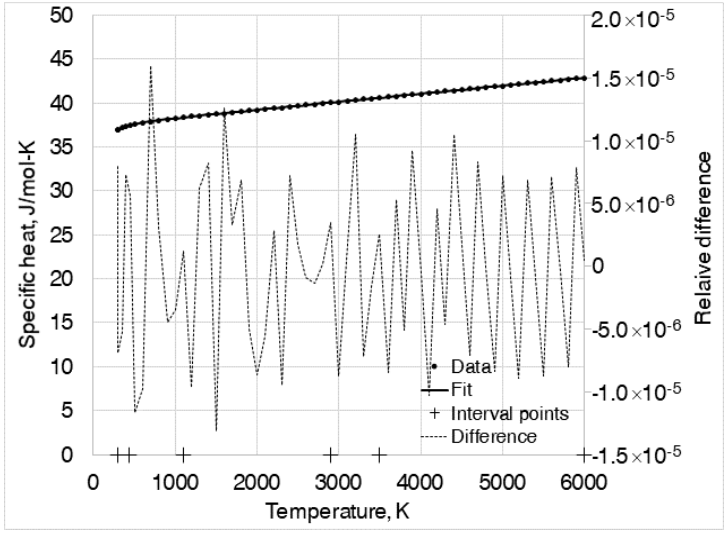


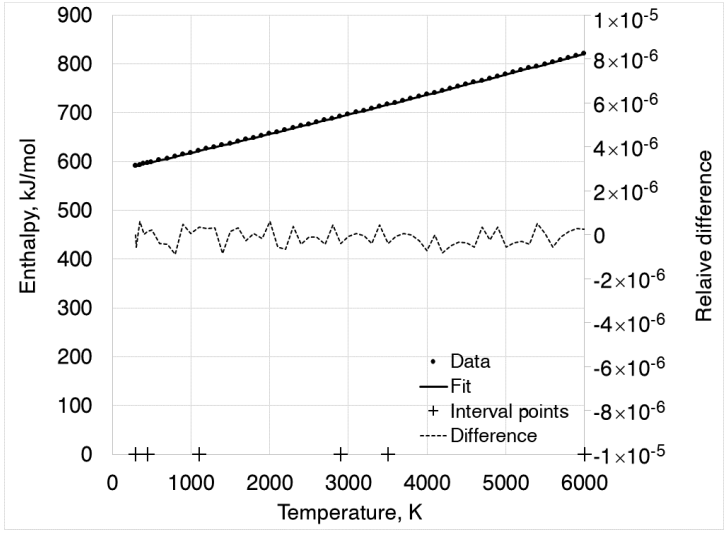


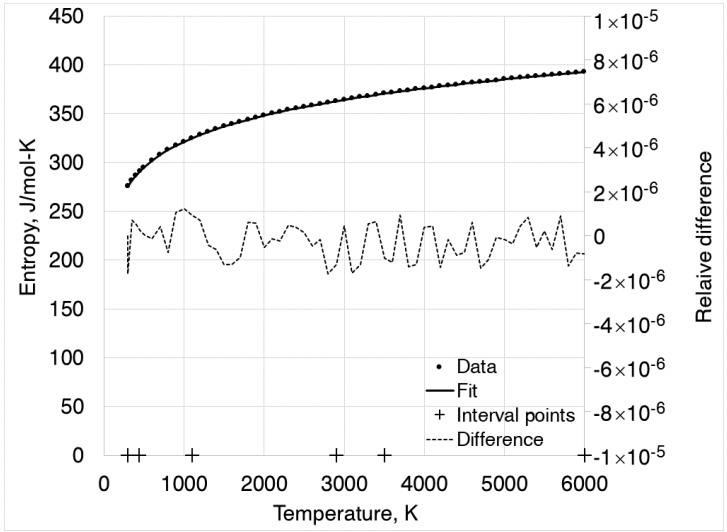


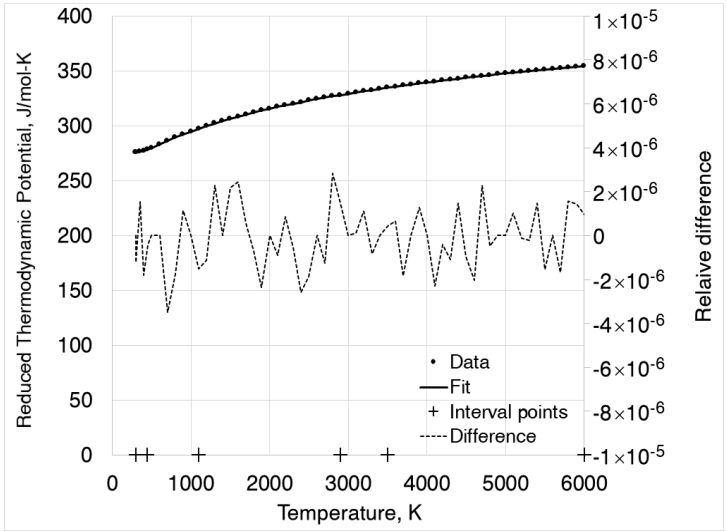
Zirconium Iodide (ZrI) – I1Zr1(g)

MW = 218.12 g/mol – $\Delta_f H^o$(298.15 K) = 591.199 kJ/mol – S^o(298.15 K) = 275.805 J/mol-K – p^o = 0.1 MPa



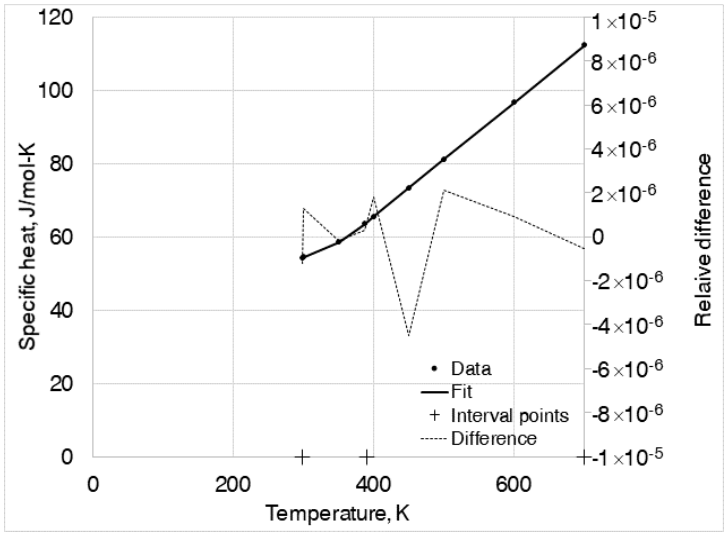


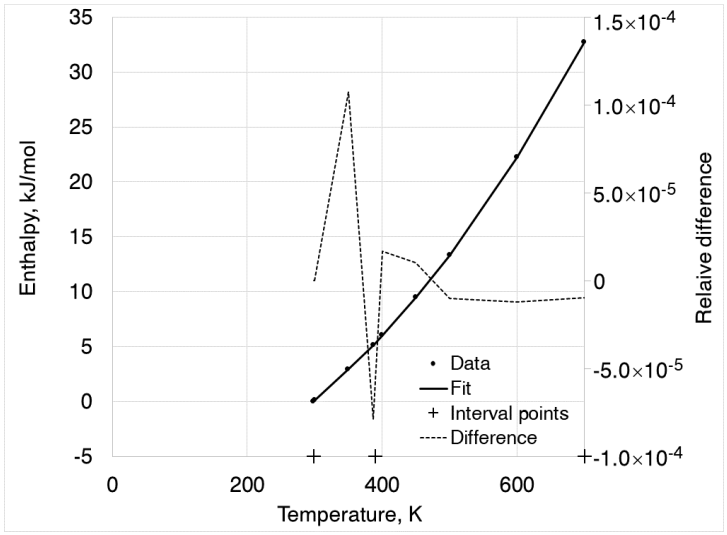


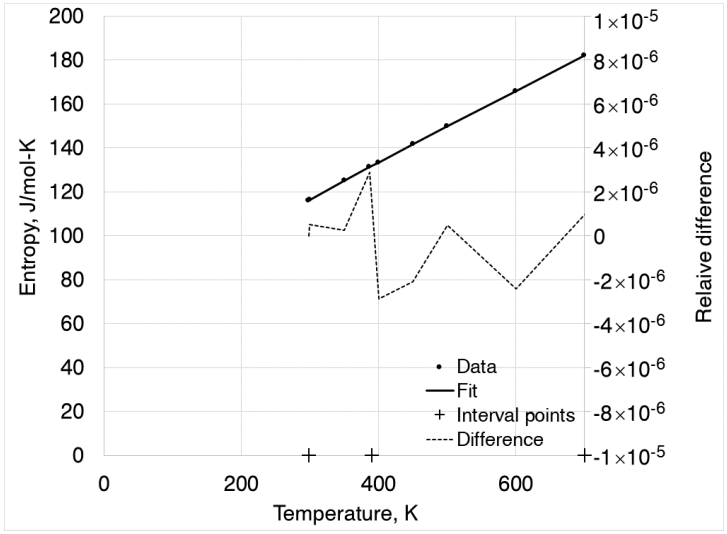


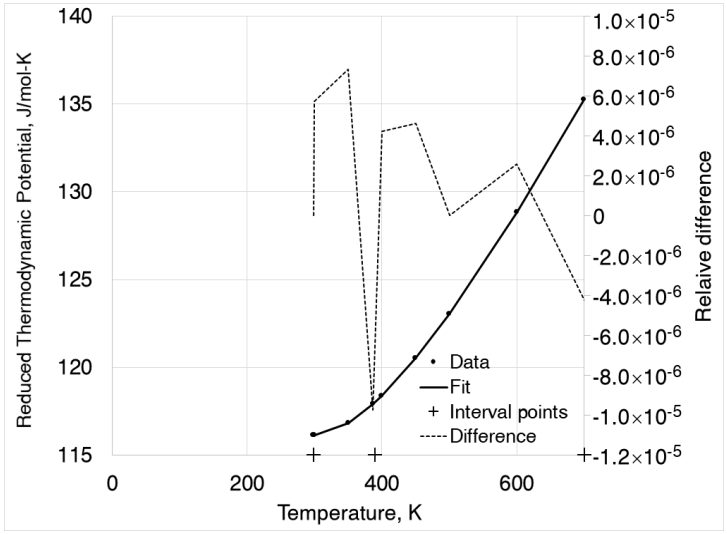
Iodine (I_2) – I2(cr)

MW = 253.809 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 116.142 J/mol-K – p^o = 0.1 MPa



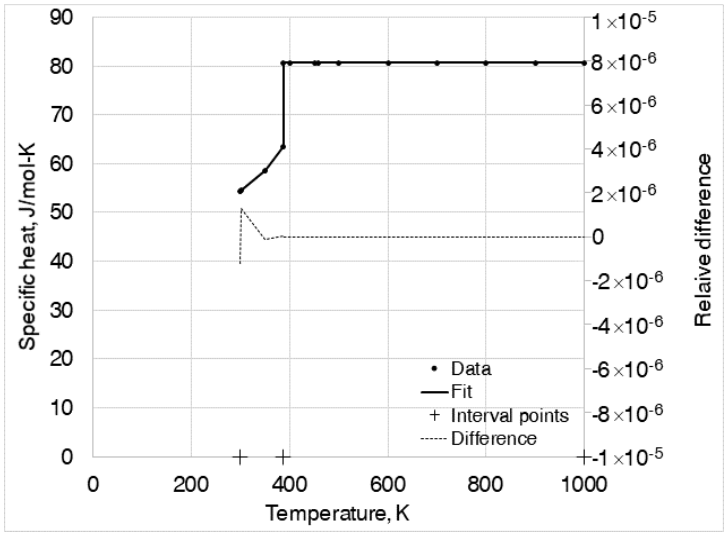


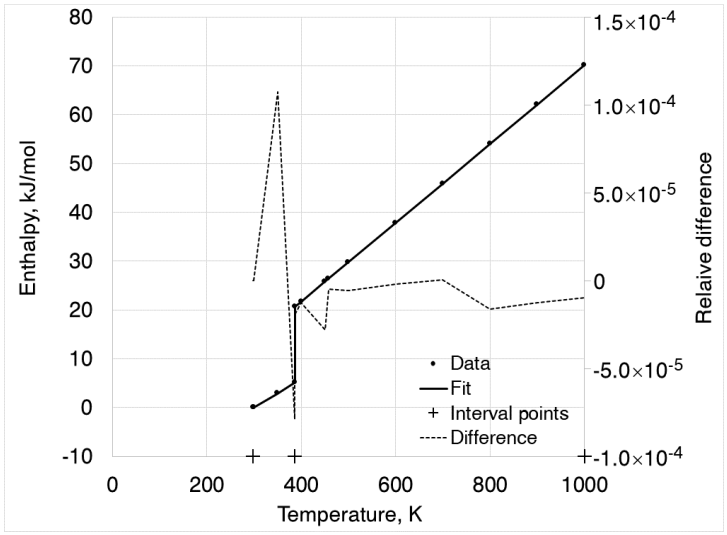


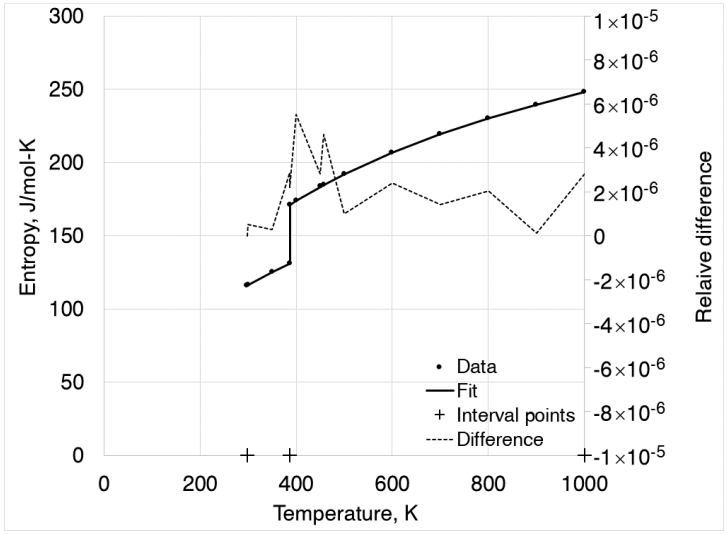


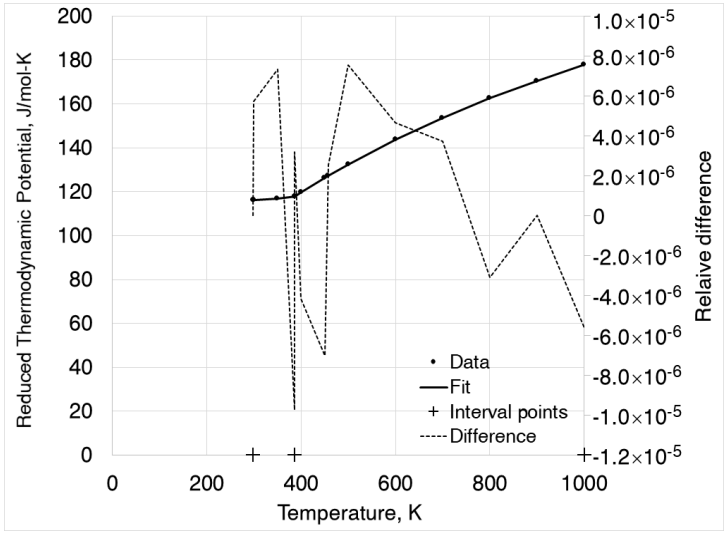
Iodine (I_2) – I2(cr,l)

MW = 253.809 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 116.142 J/mol-K – p^o = 0.1 MPa



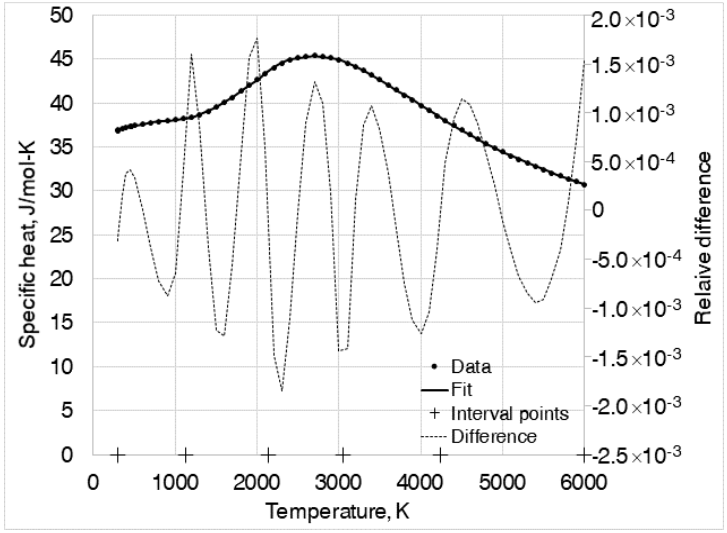


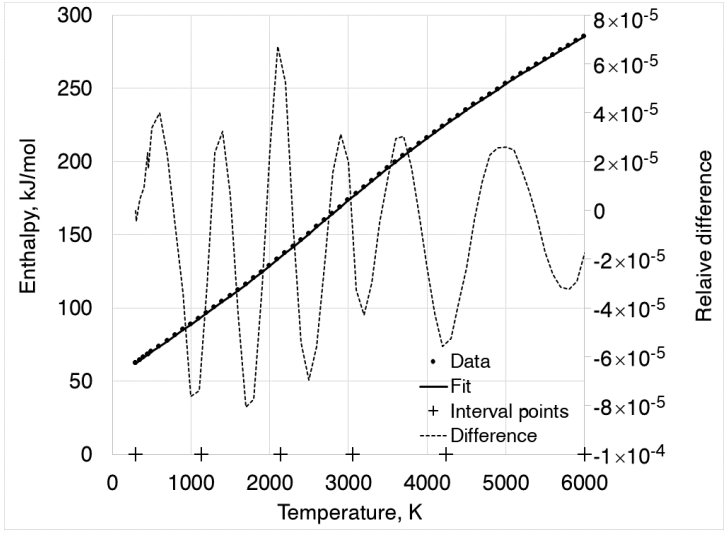


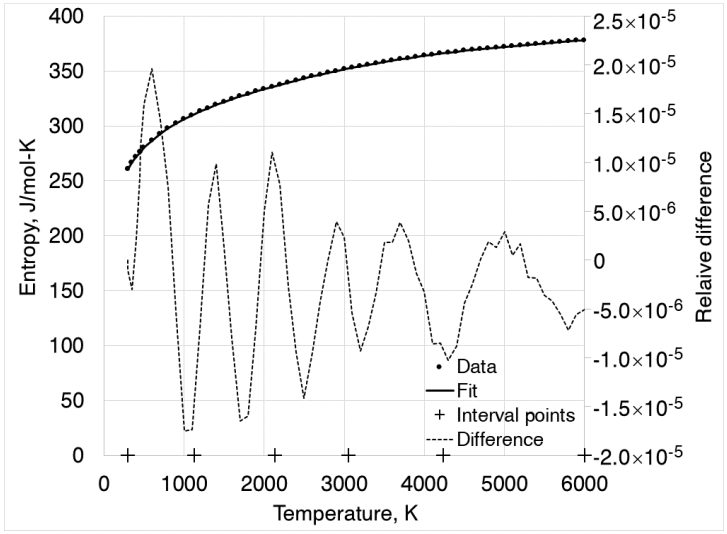


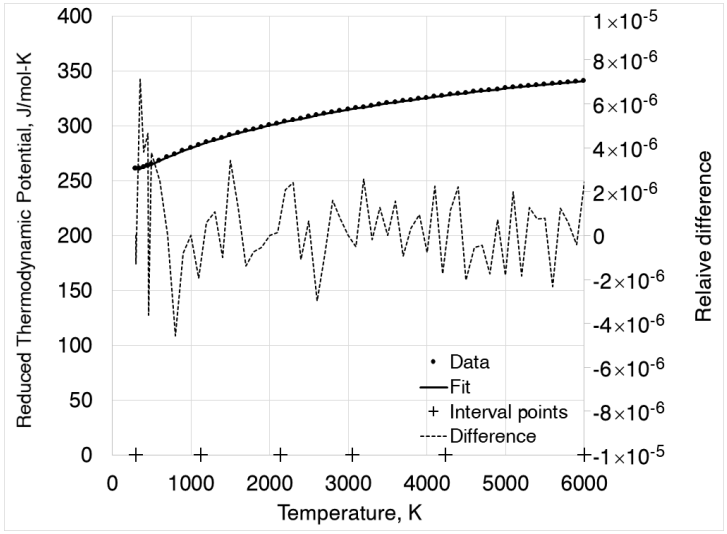
Iodine (I_2) – I2(g)

MW = 253.809 g/mol – $\Delta_f H^o$(298.15 K) = 62.421 kJ/mol – S^o(298.15 K) = 260.685 J/mol-K – p^o = 0.1 MPa



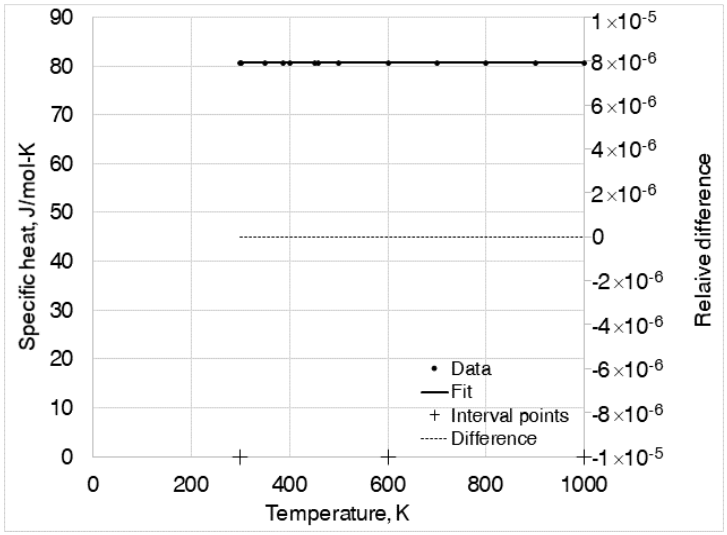


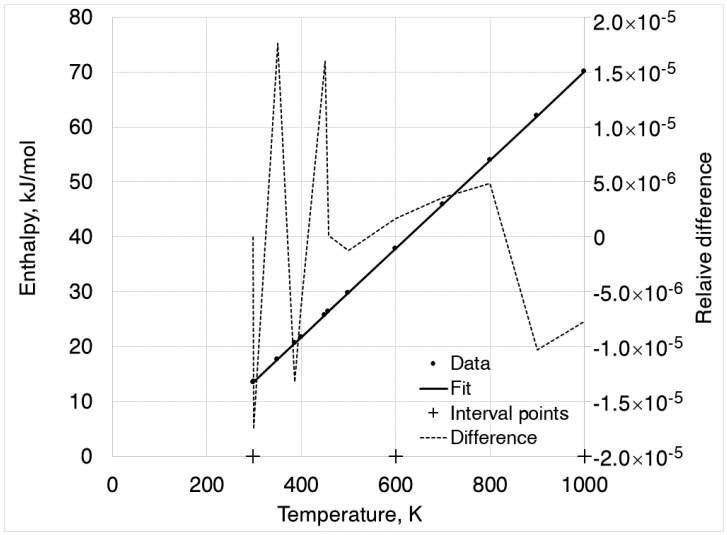


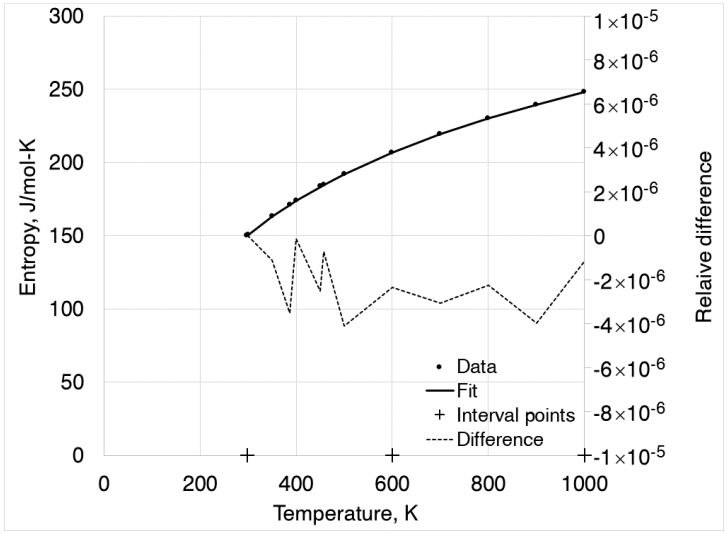


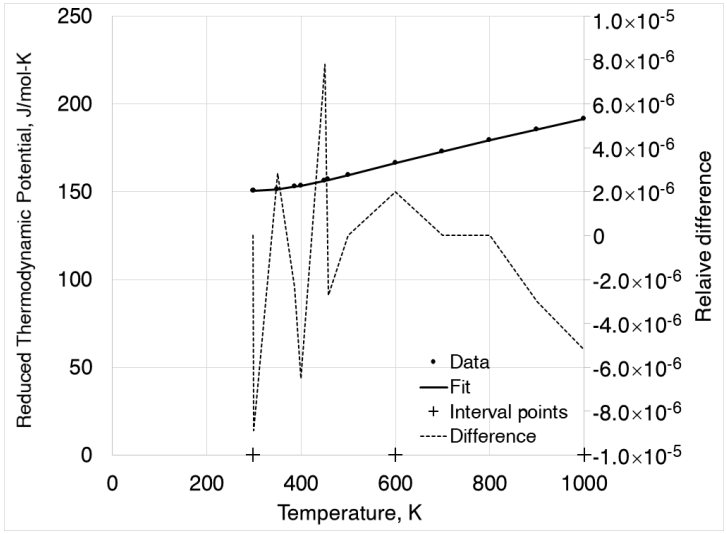
Iodine (I_2) – I2(l)

MW = 253.809 g/mol – $\Delta_f H^o$(298.15 K) = 13.523 kJ/mol – S^o(298.15 K) = 150.356 J/mol-K – p^o = 0.1 MPa



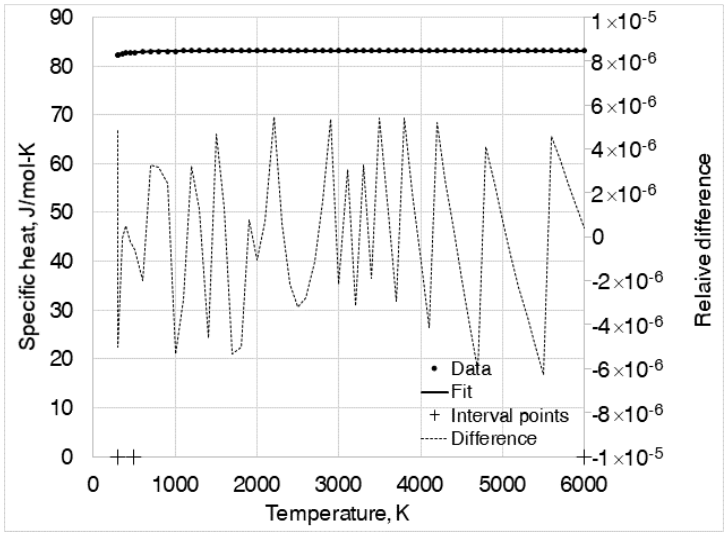


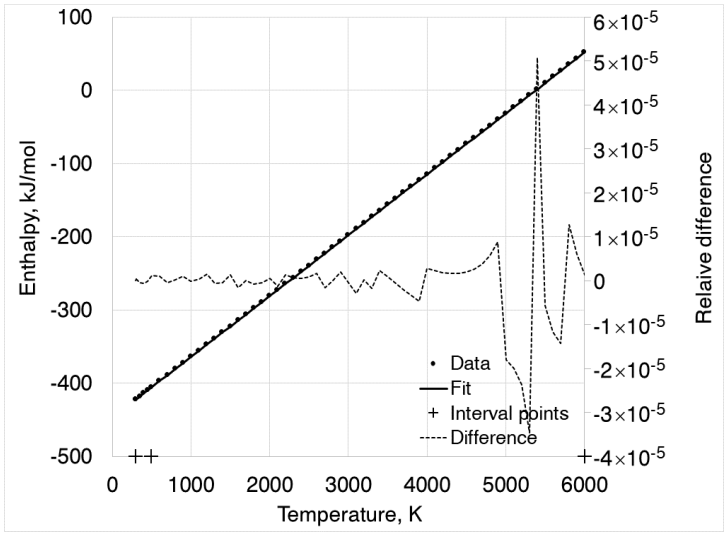


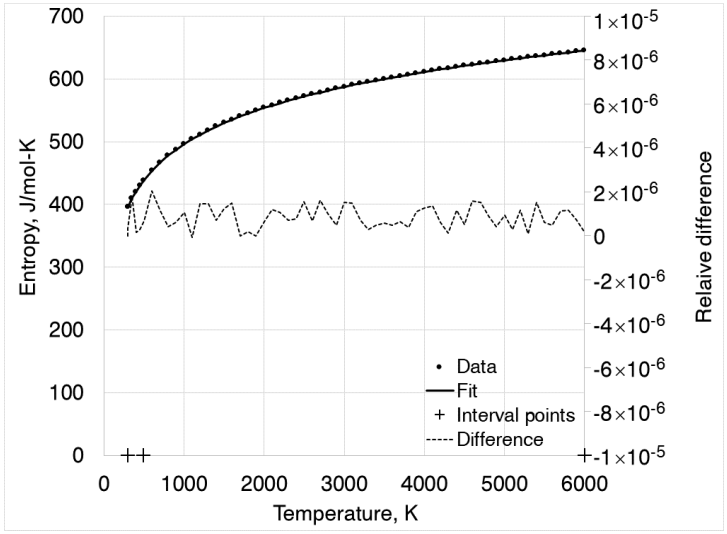


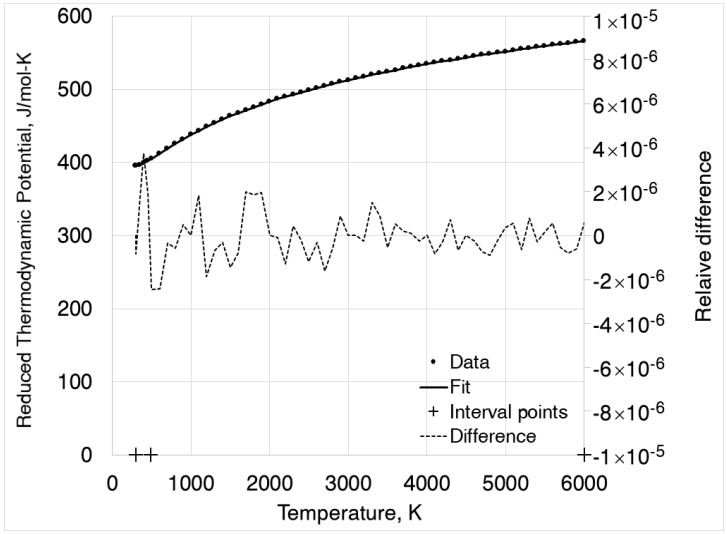
Potassium Iodide ($(KI)_2$) – I2K2(g)

MW = 332.0056 g/mol – $\Delta_f H^o$(298.15 K) = -422.166 kJ/mol – S^o(298.15 K) = 395.772 J/mol-K – p^o = 0.1 MPa



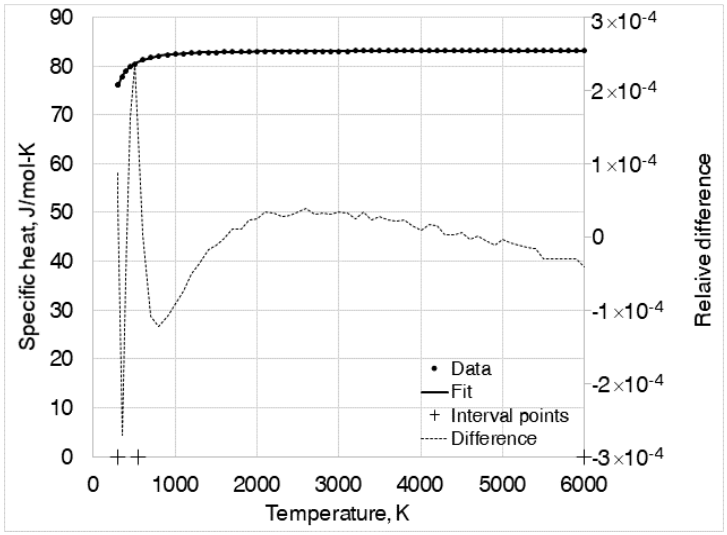


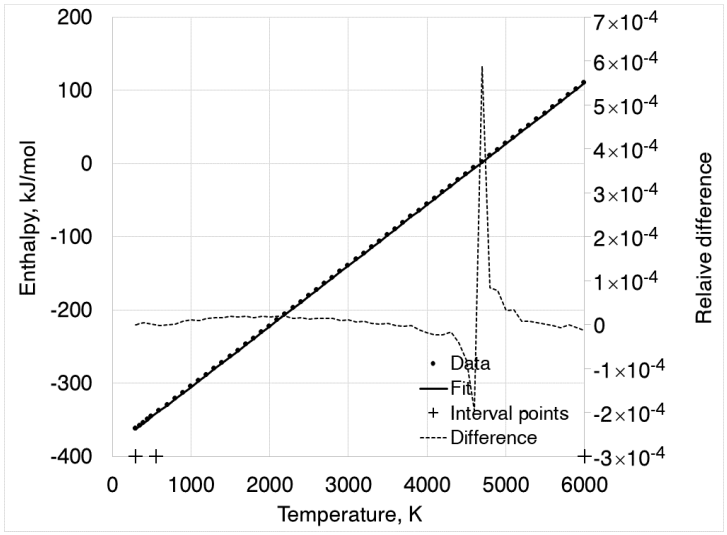


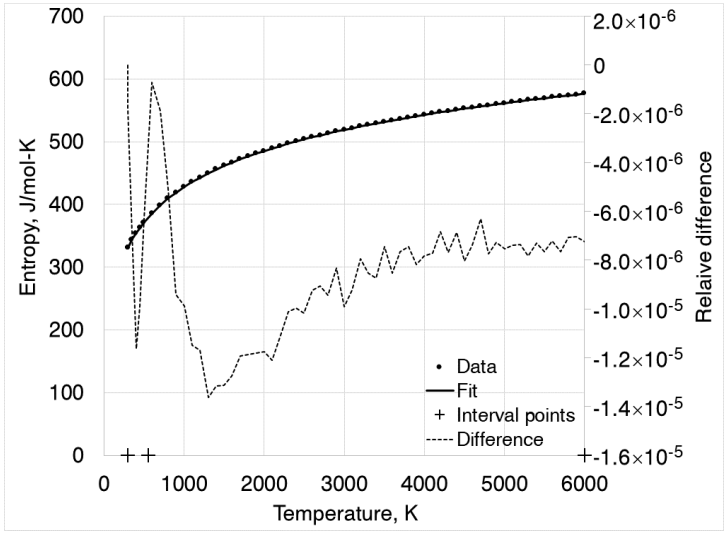


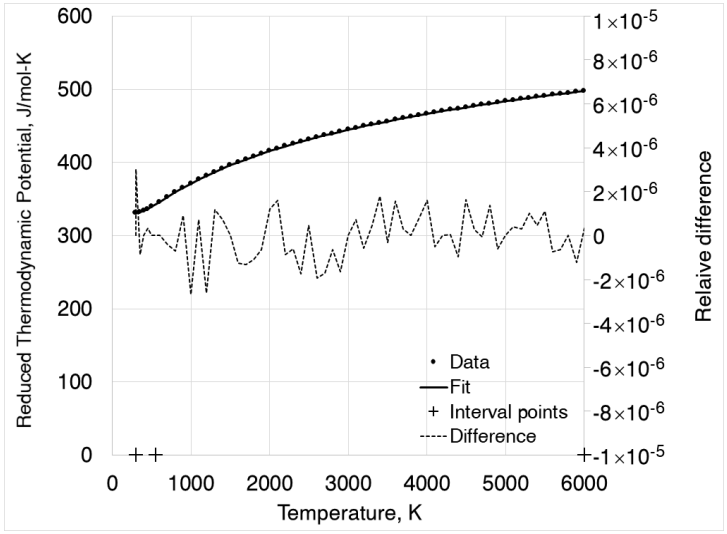
Lithium Iodide ($(LiI)_2$) – I2Li2(g)

MW = 267.691 g/mol – $\Delta_f H^o$(298.15 K) = -361.916 kJ/mol – S^o(298.15 K) = 330.628 J/mol-K – p^o = 0.1 MPa



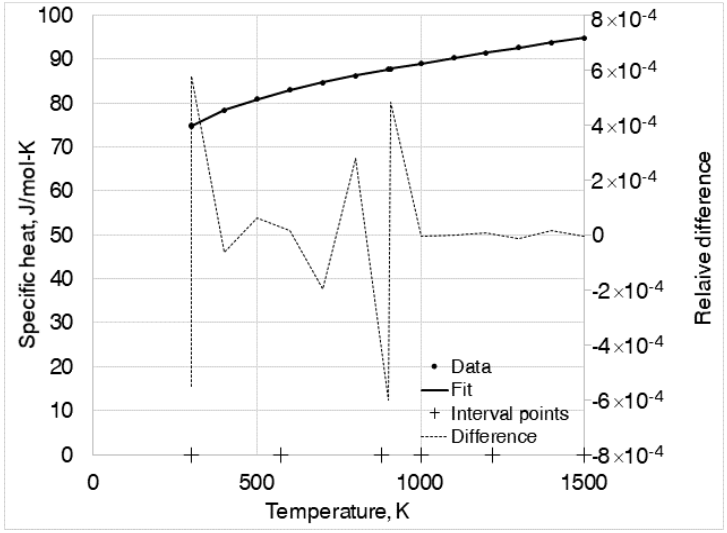


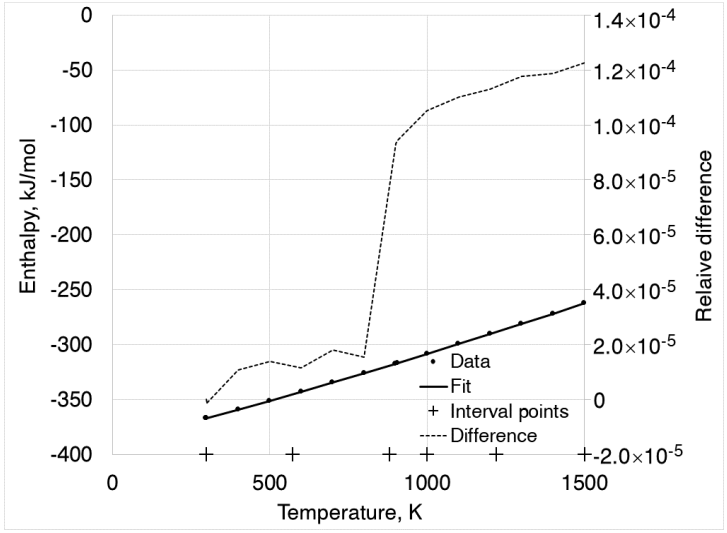


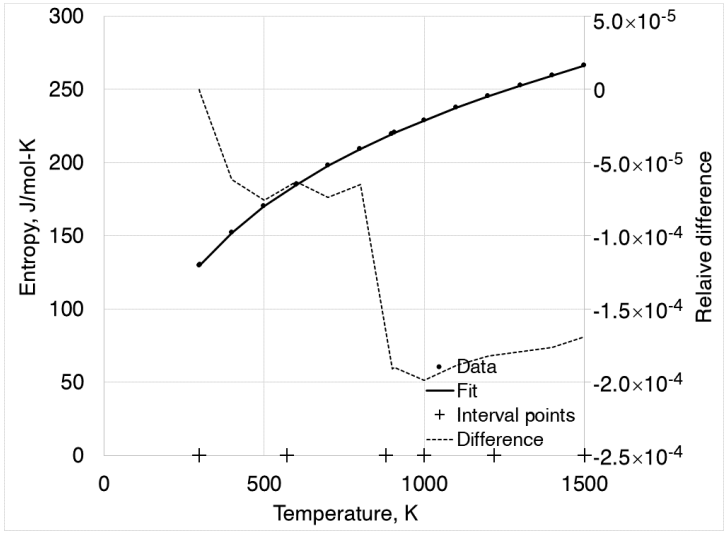


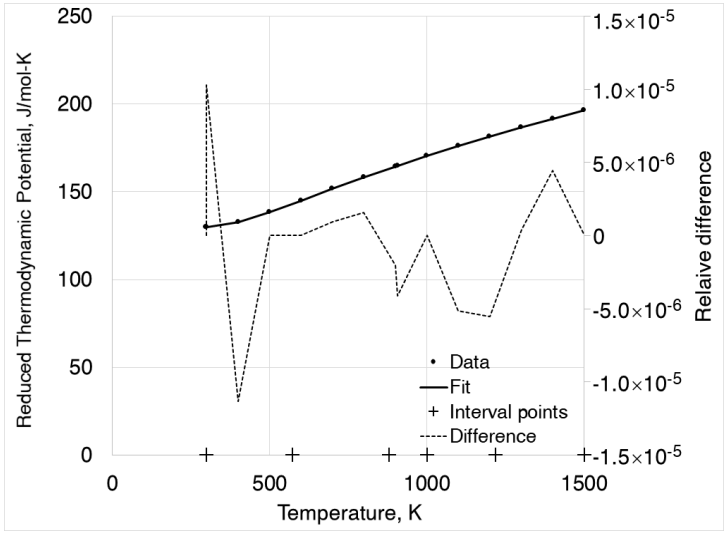
Magnesium Iodide (MgI_2) – I2Mg1(cr)

MW = 278.114 g/mol – $\Delta_f H^o$(298.15 K) = -366.937 kJ/mol – S^o(298.15 K) = 129.704 J/mol-K – p^o = 0.1 MPa



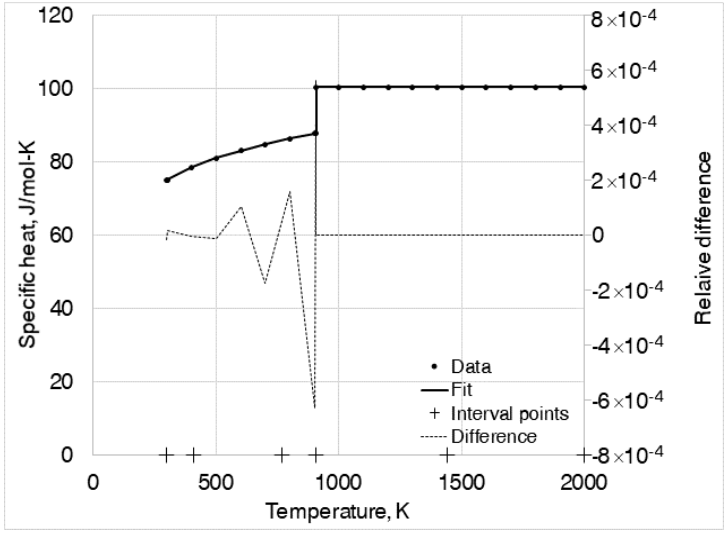


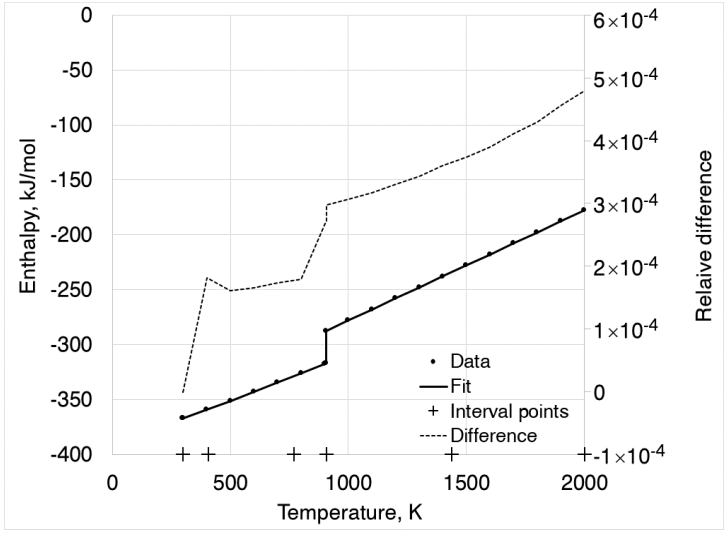


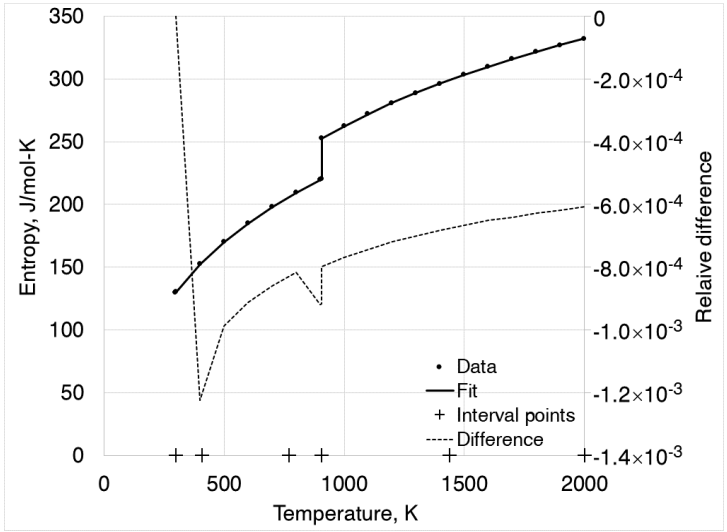


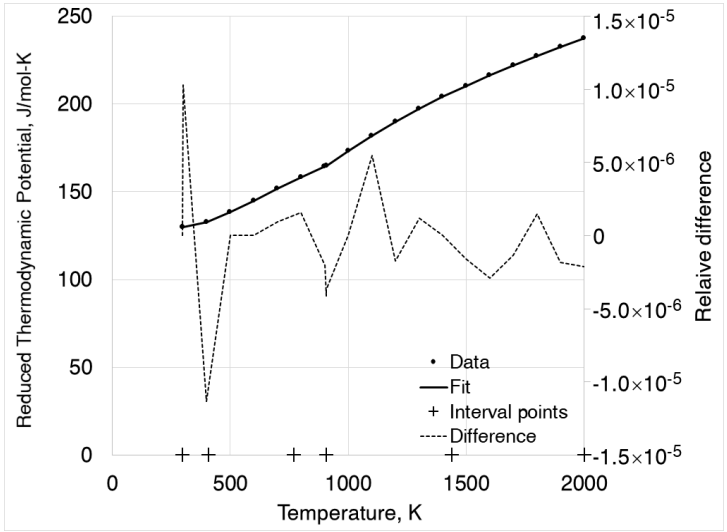
Magnesium Iodide (MgI_2) – I2Mg1(cr,l)

MW = 278.114 g/mol – $\Delta_f H^o$(298.15 K) = -366.937 kJ/mol – S^o(298.15 K) = 129.704 J/mol-K – p^o = 0.1 MPa



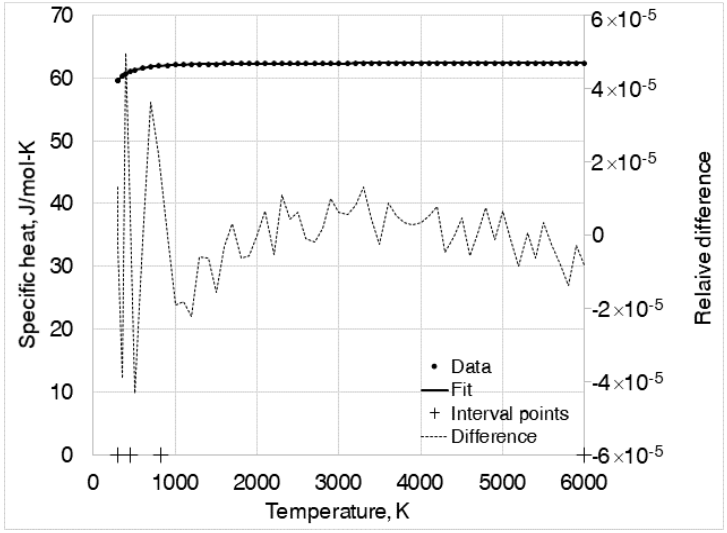


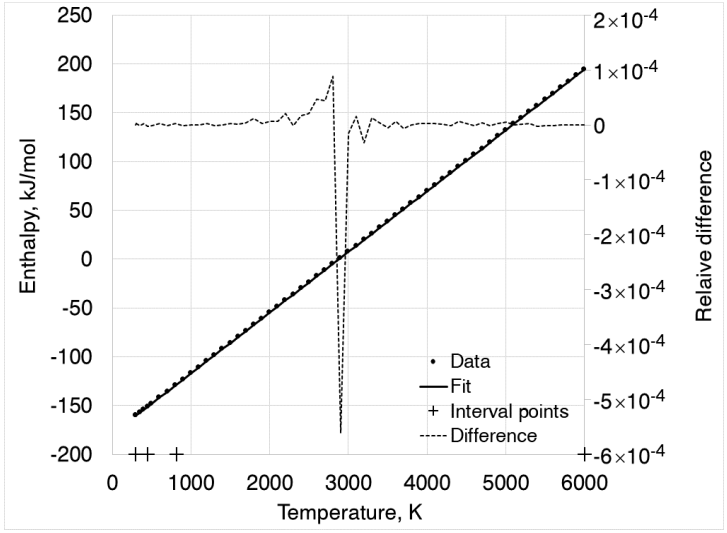


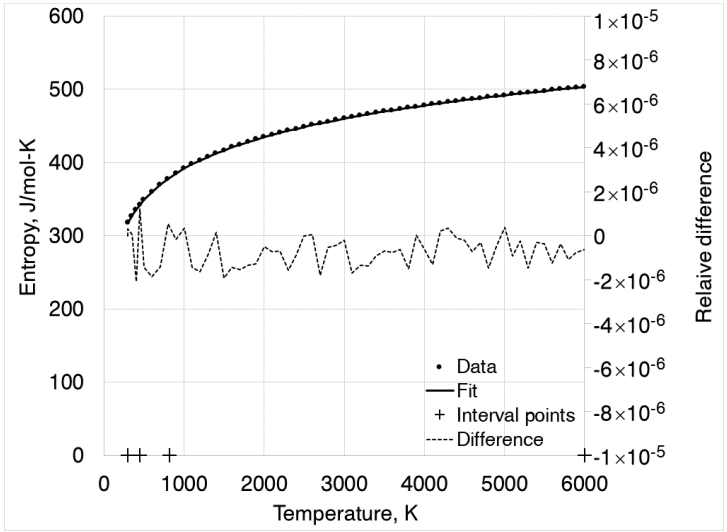


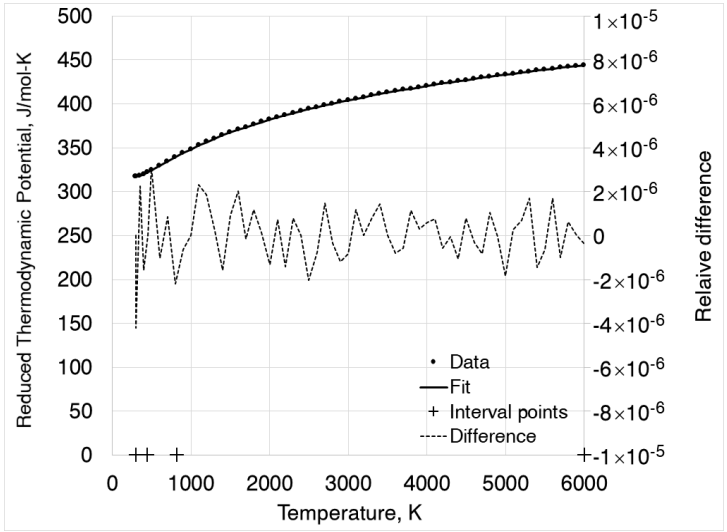
Magnesium Iodide (MgI_2) – I2Mg1(g)

MW = 278.114 g/mol – $\Delta_f H^o$(298.15 K) = -160.247 kJ/mol – S^o(298.15 K) = 317.489 J/mol-K – p^o = 0.1 MPa



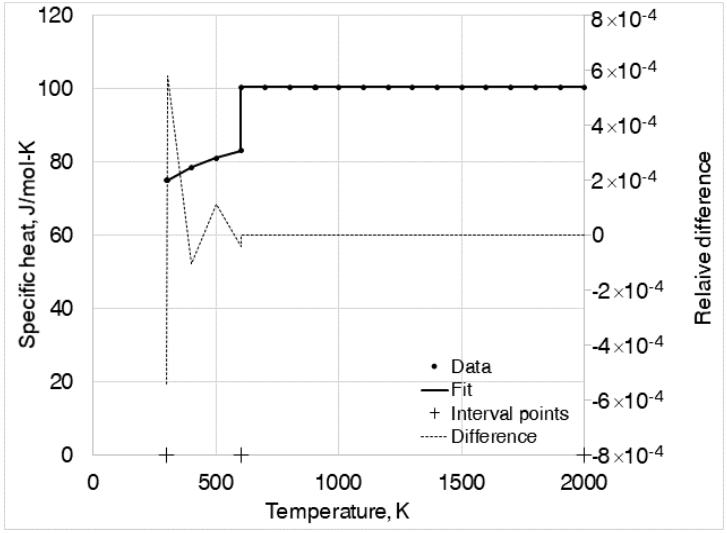


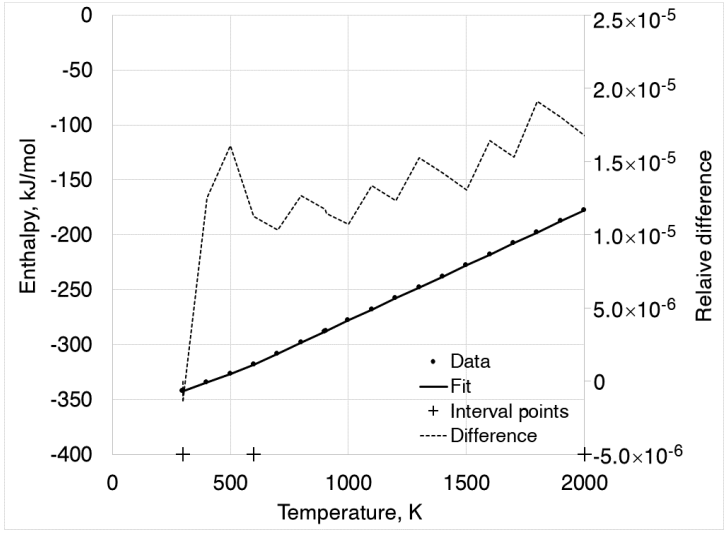


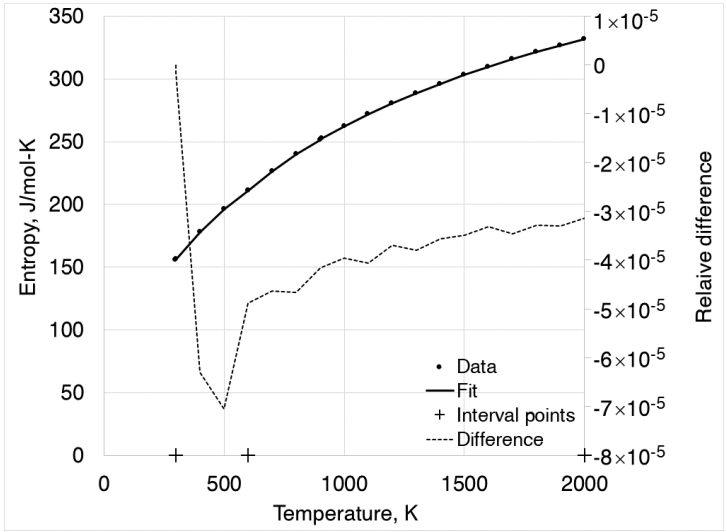


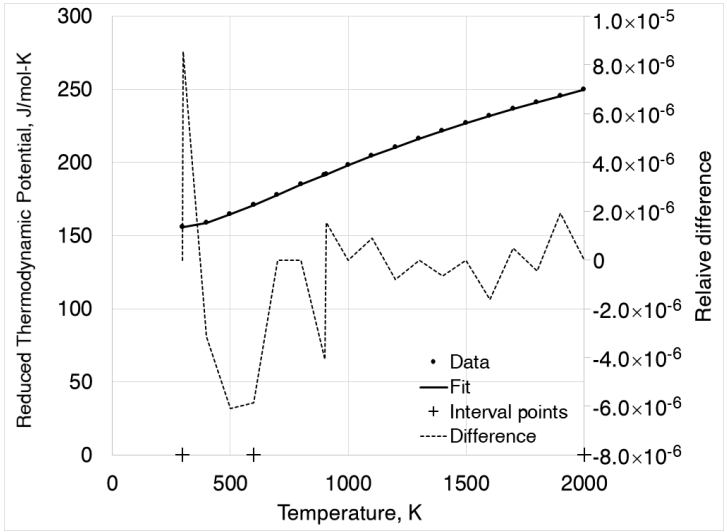
Magnesium Iodide (MgI_2) – I2Mg1(l)

MW = 278.114 g/mol – $\Delta_f H^o$(298.15 K) = -342.252 kJ/mol – S^o(298.15 K) = 155.737 J/mol-K – p^o = 0.1 MPa



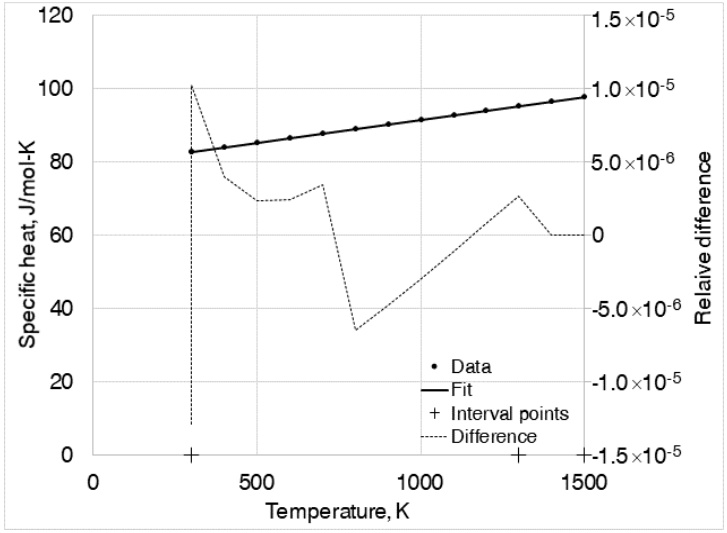


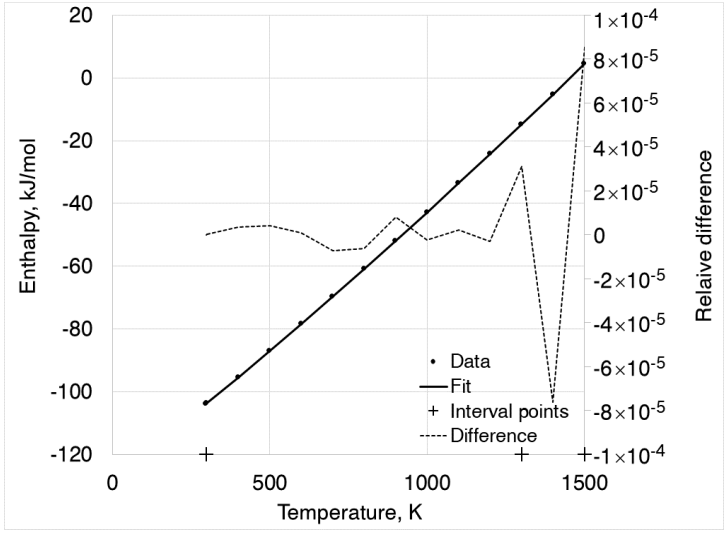


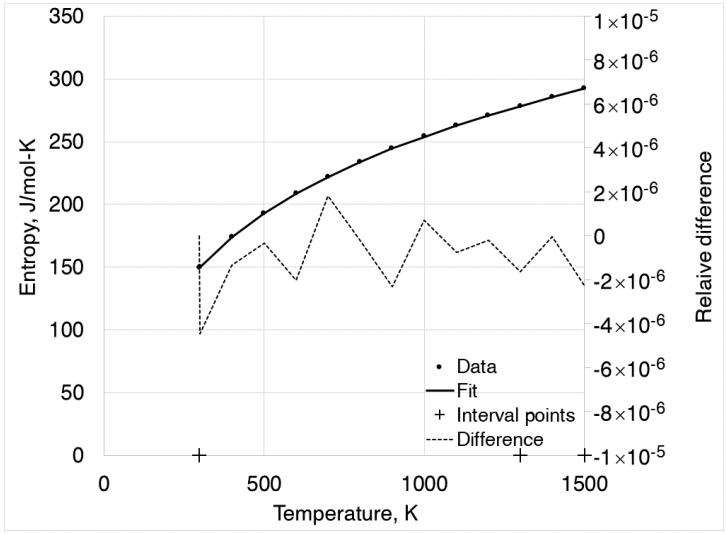


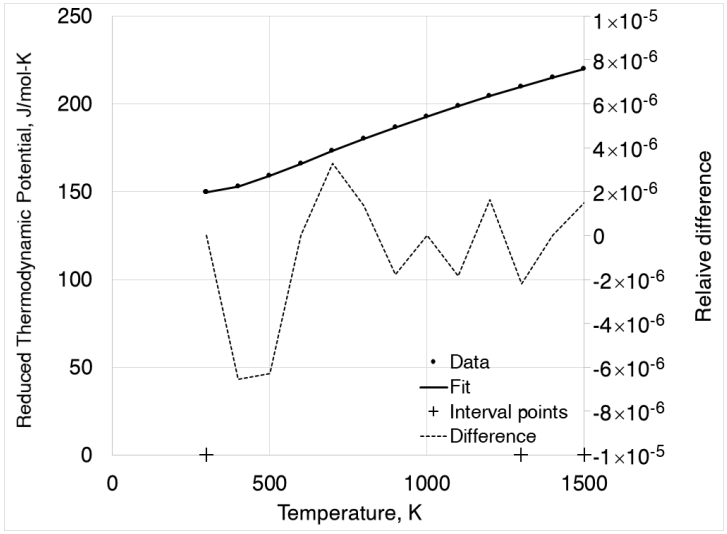
Molybdenum Iodide, Alpha (MoI_2) – I2Mo1(cr)

MW = 349.75 g/mol – $\Delta_f H^o$(298.15 K) = -103.931 kJ/mol – S^o(298.15 K) = 149.662 J/mol-K – p^o = 0.1 MPa



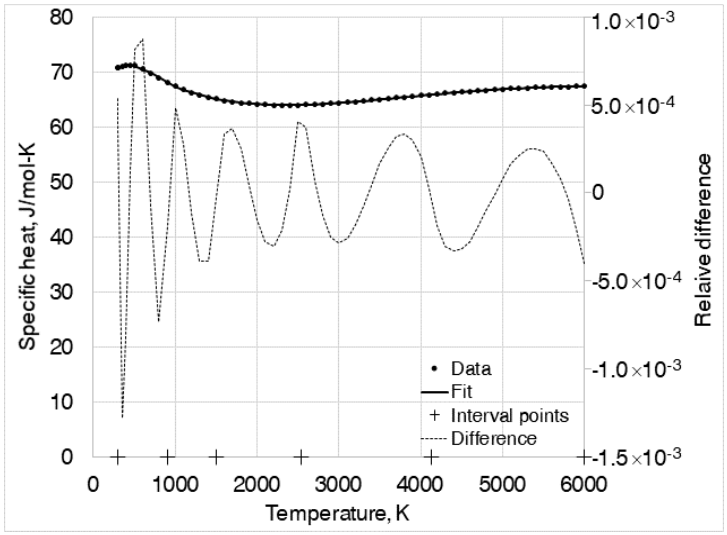


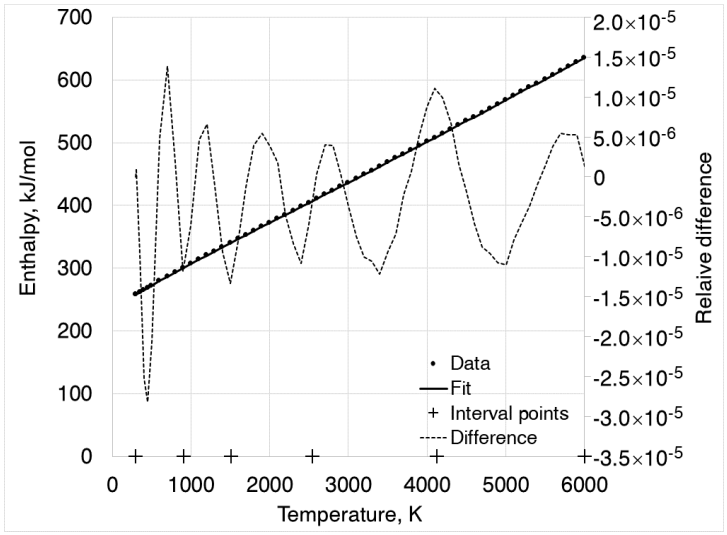


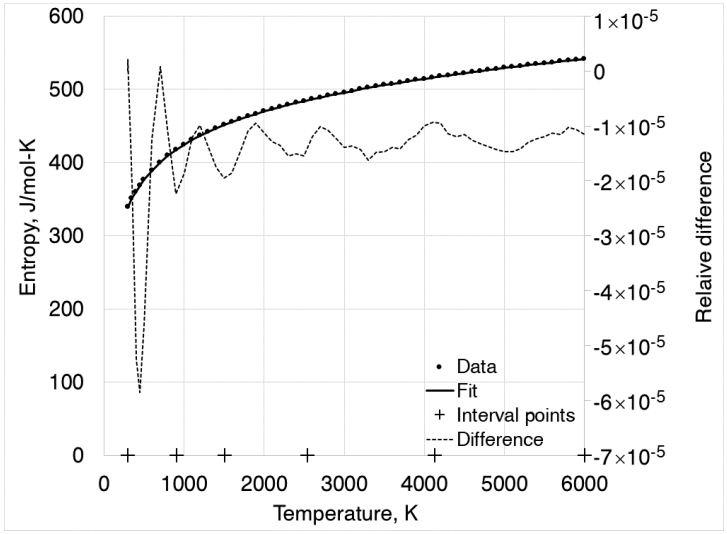


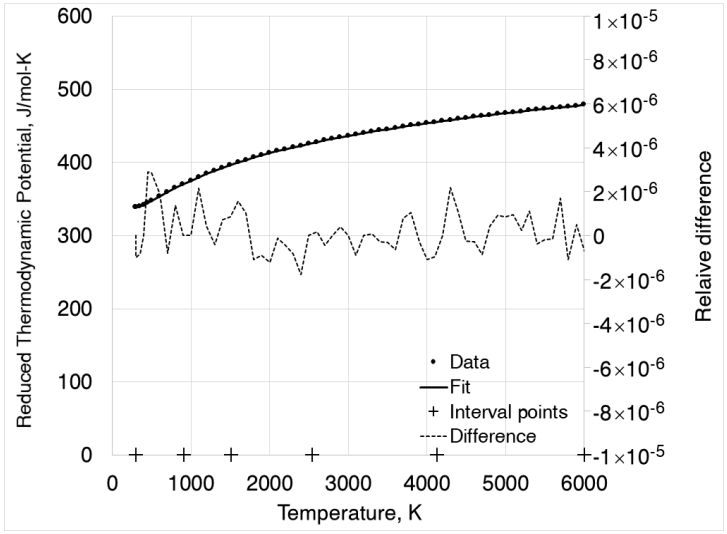
Molybdenum Iodide (MoI_2) – I2Mo1(g)

MW = 349.75 g/mol – $\Delta_f H^o$(298.15 K) = 257.734 kJ/mol – S^o(298.15 K) = 339.29 J/mol-K – p^o = 0.1 MPa



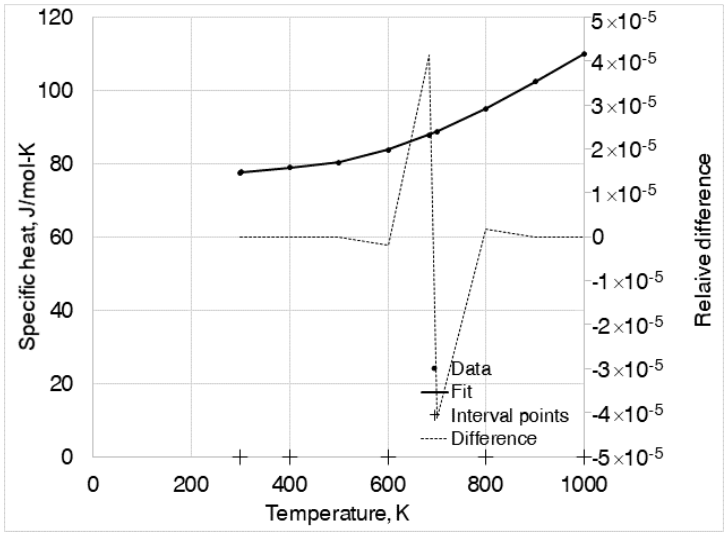


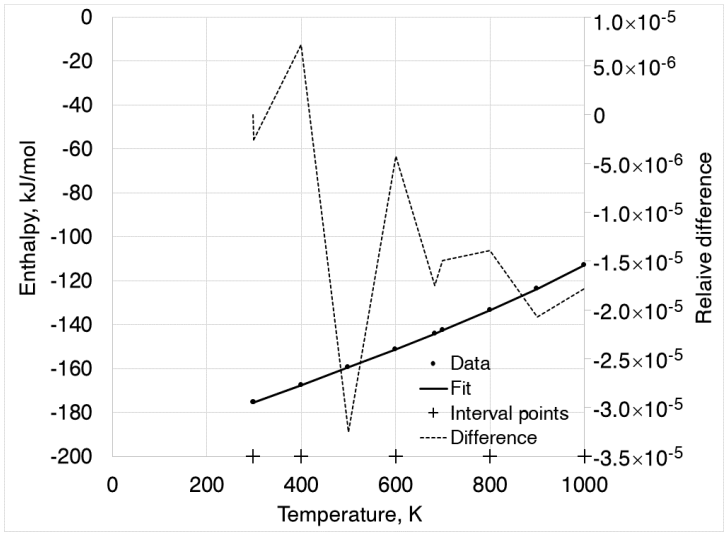


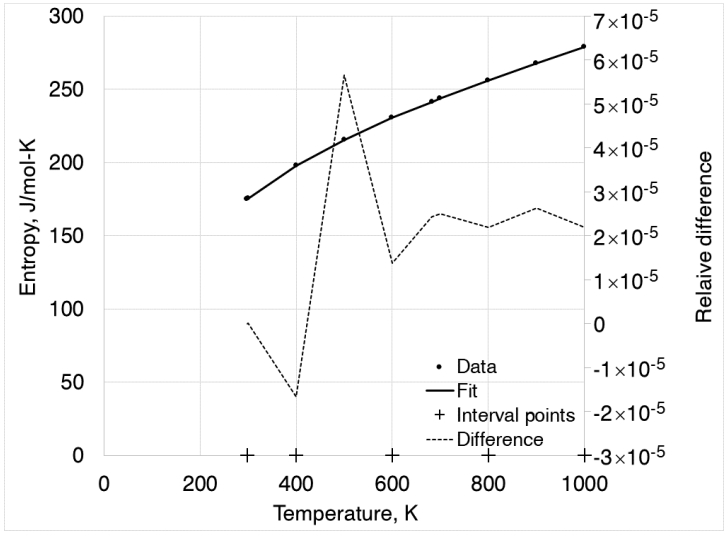


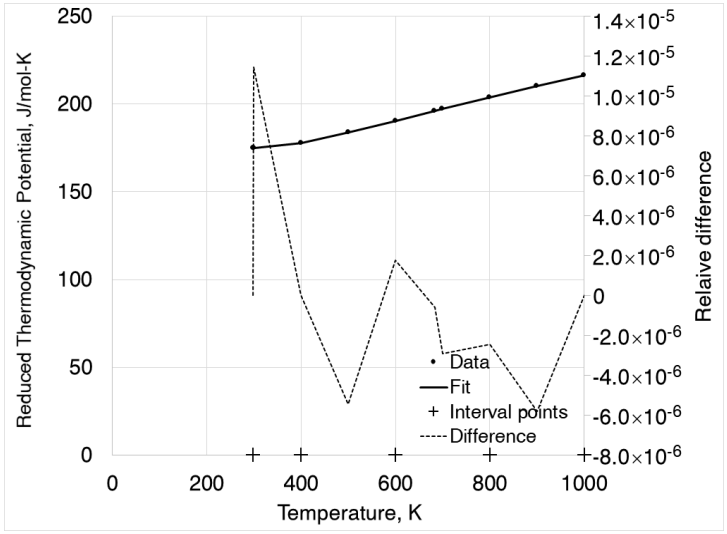
Lead Iodide (PbI_2) – I2Pb1(cr)

MW = 461 g/mol – $\Delta_f H^o$(298.15 K) = -175.393 kJ/mol – S^o(298.15 K) = 174.836 J/mol-K – p^o = 0.1 MPa



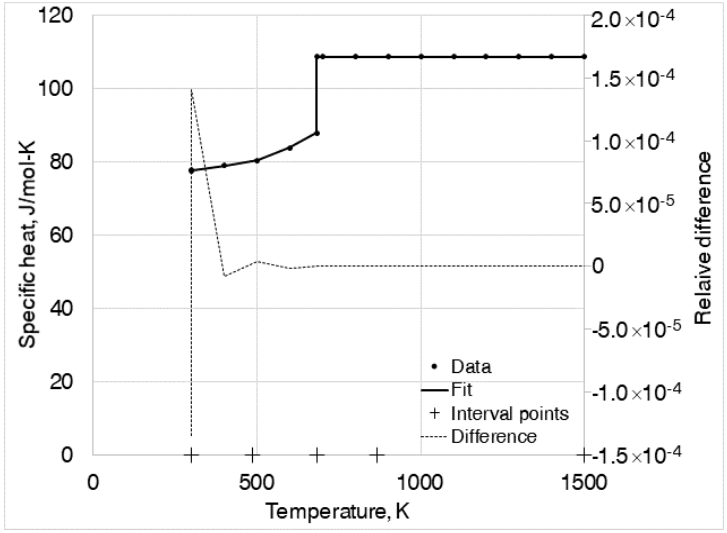


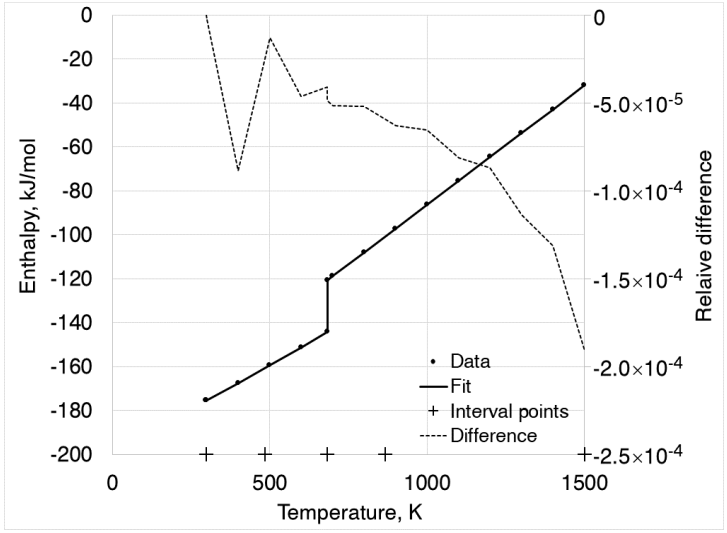


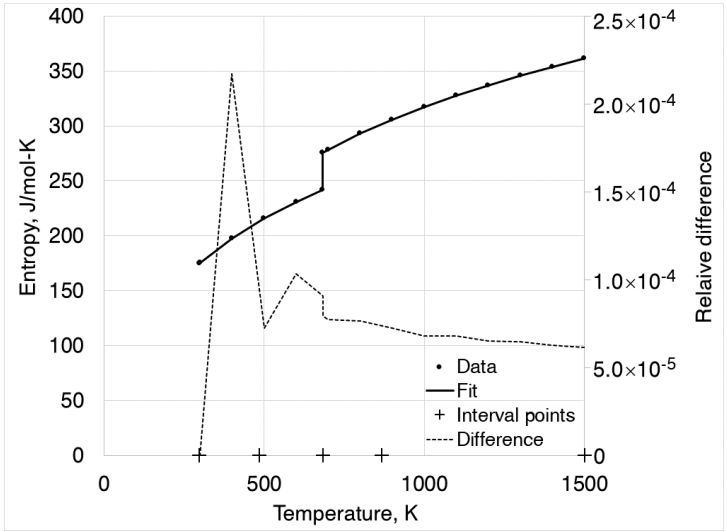


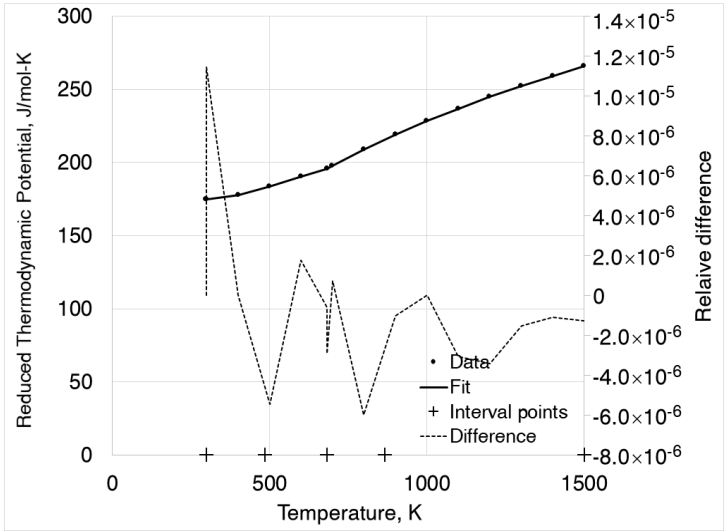
Lead Iodide (PbI_2) – I2Pb1(cr,l)

MW = 461 g/mol – $\Delta_f H^o$(298.15 K) = -175.393 kJ/mol – S^o(298.15 K) = 174.836 J/mol-K – p^o = 0.1 MPa



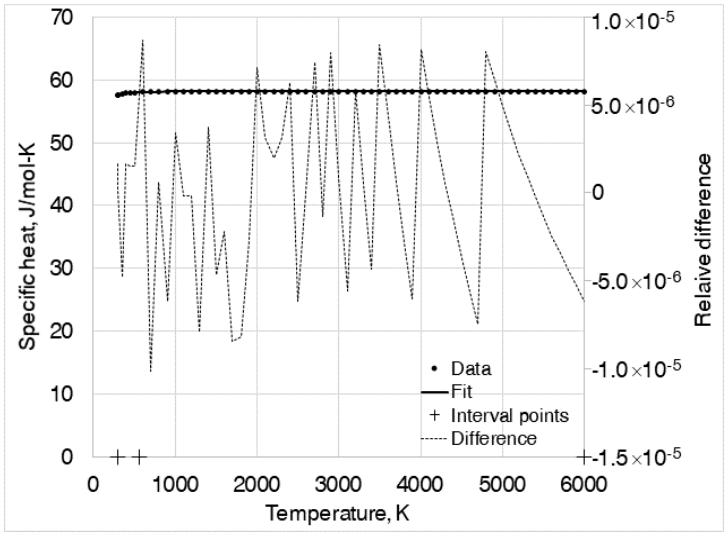


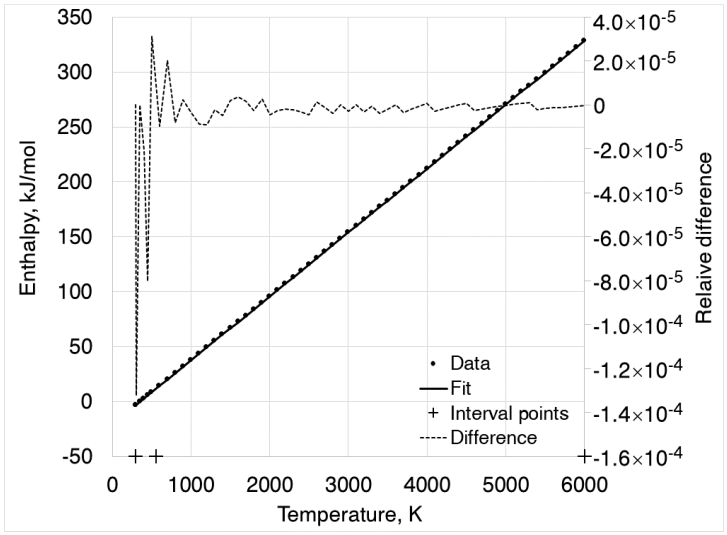


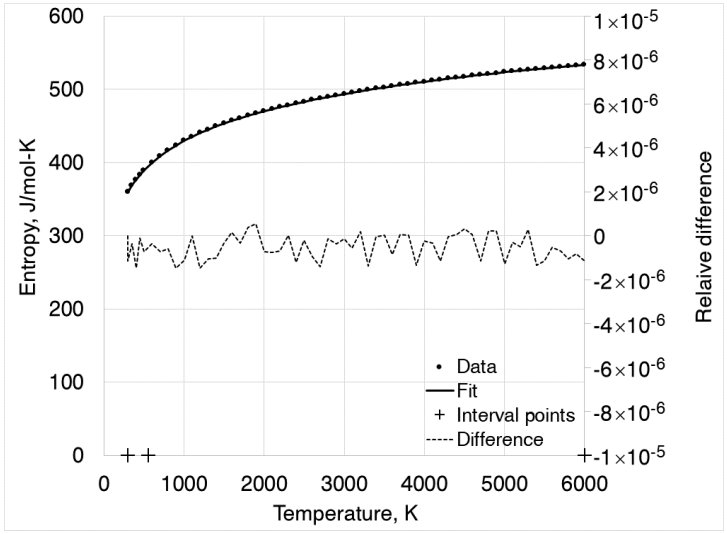


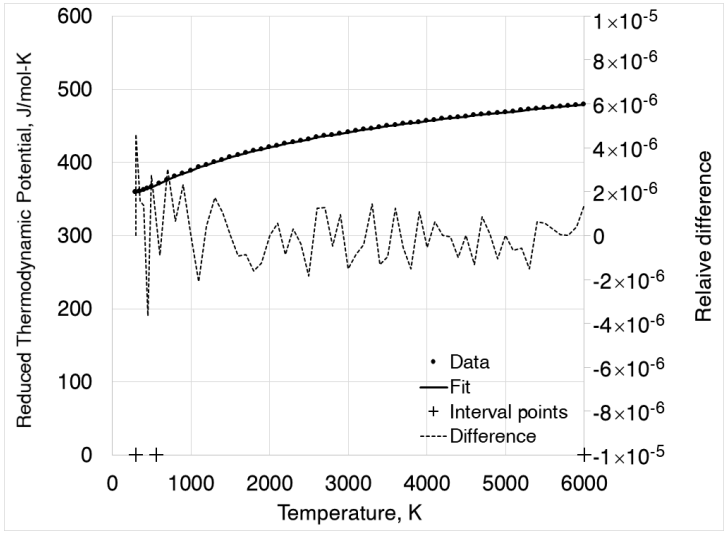
Lead Iodide (PbI_2) – I2Pb1(g)

MW = 461 g/mol – $\Delta_f H^o$(298.15 K) = -3.18 kJ/mol – S^o(298.15 K) = 359.552 J/mol-K – p^o = 0.1 MPa



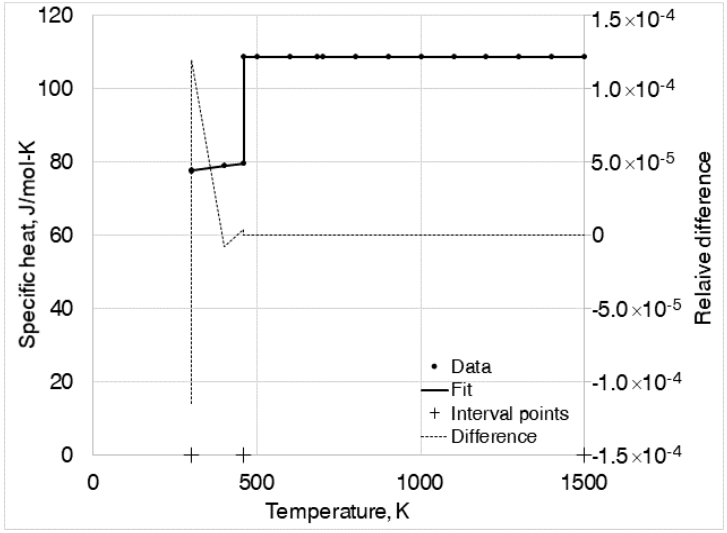


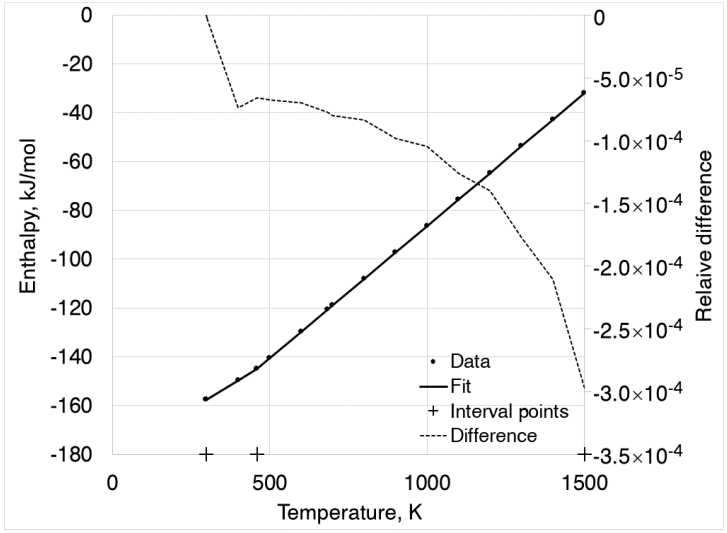


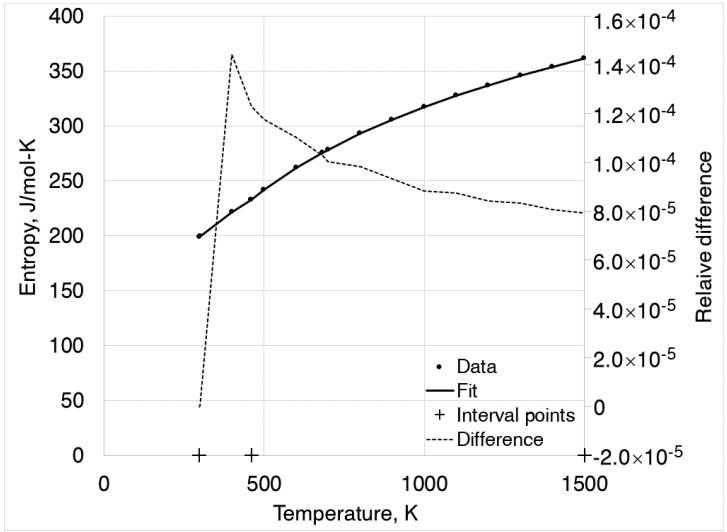


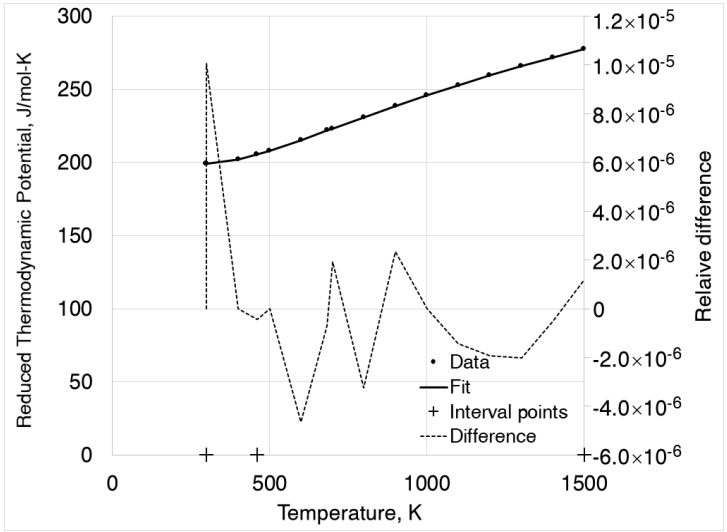
Lead Iodide (PbI_2) – I2Pb1(l)

MW = 461 g/mol – $\Delta_f H^o$(298.15 K) = -157.679 kJ/mol – S^o(298.15 K) = 198.902 J/mol-K – p^o = 0.1 MPa



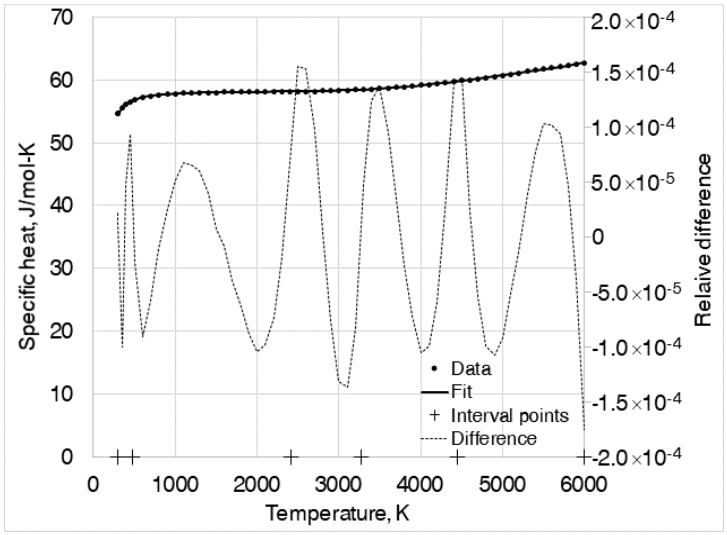


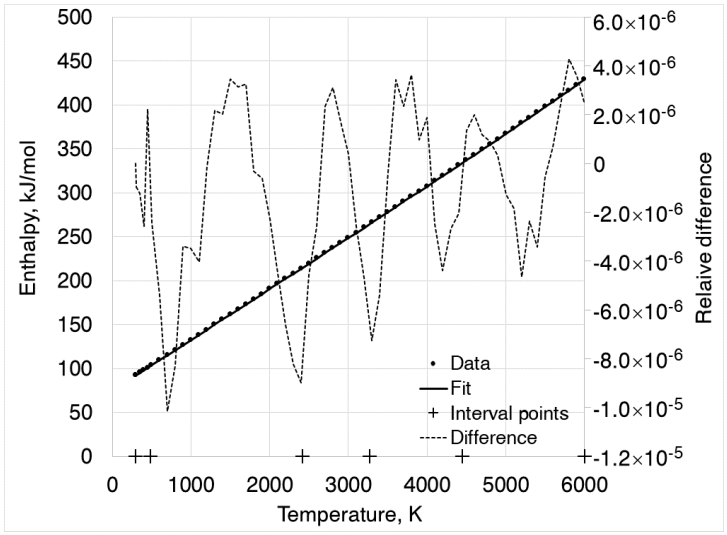


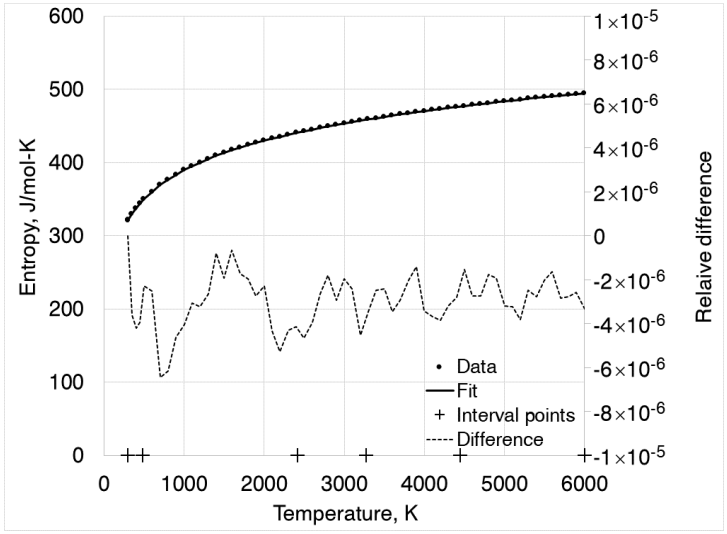


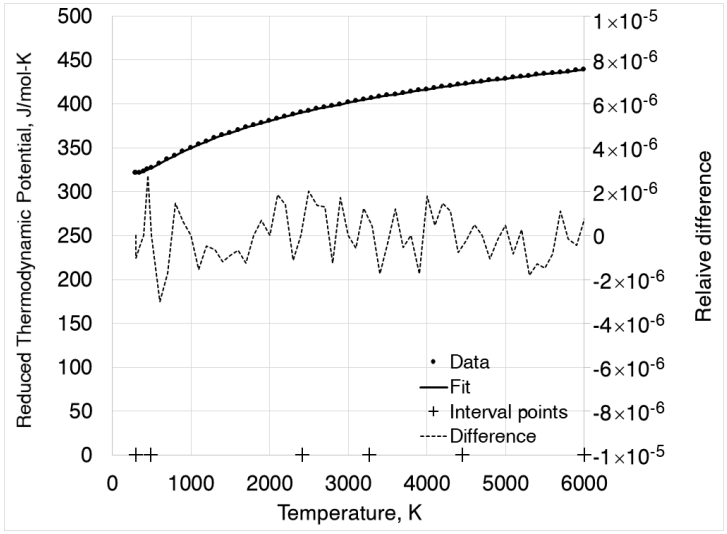
Diiodosilylene (SiI_2) – I2Si1(g)

MW = 281.8945 g/mol – $\Delta_f H^o$(298.15 K) = 92.466 kJ/mol – S^o(298.15 K) = 320.972 J/mol-K – p^o = 0.1 MPa



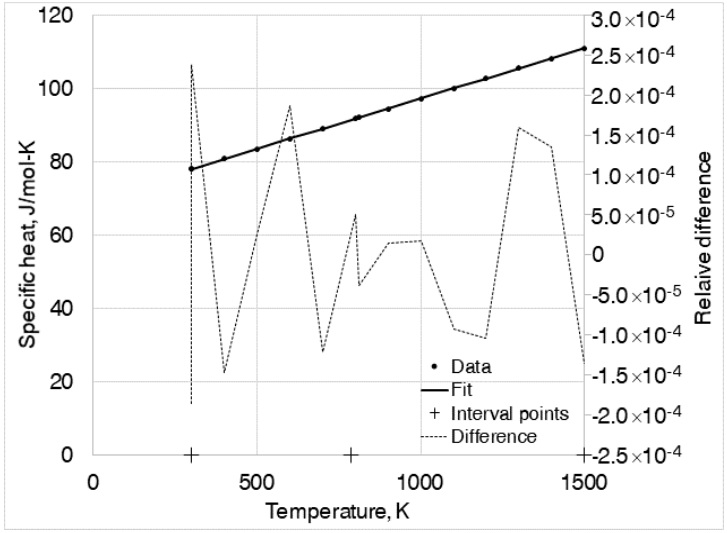


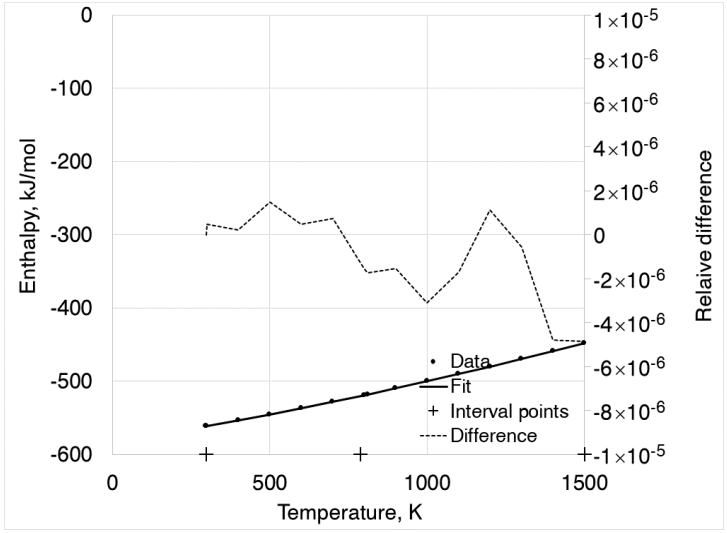


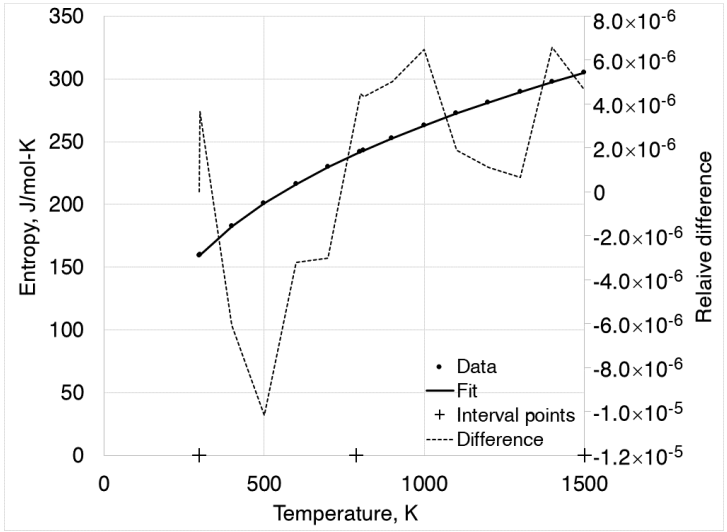


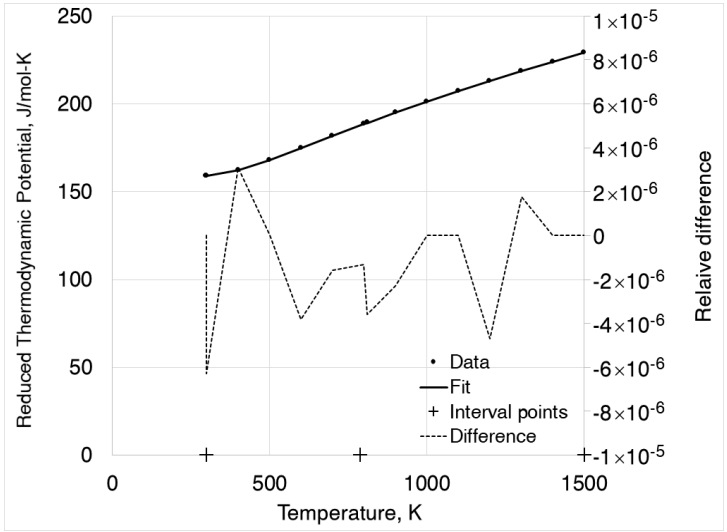
Strontium Iodide (SrI_2) – I2Sr1(cr)

MW = 341.43 g/mol – $\Delta_f H^o$(298.15 K) = -561.493 kJ/mol – S^o(298.15 K) = 159.12 J/mol-K – p^o = 0.1 MPa



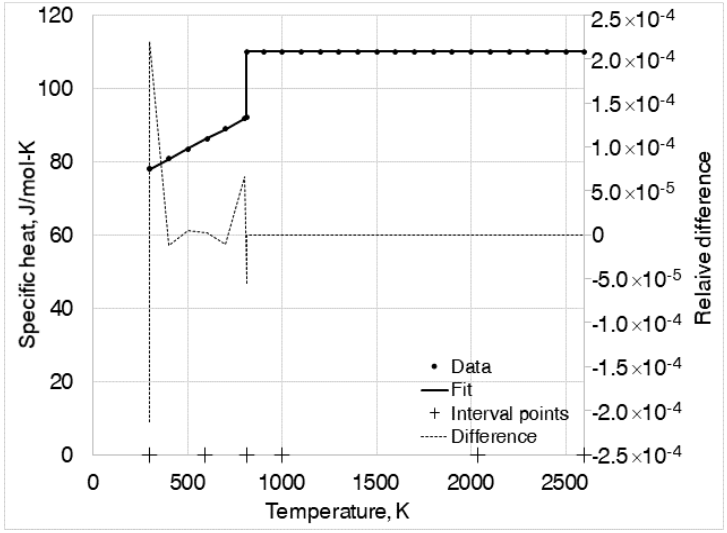


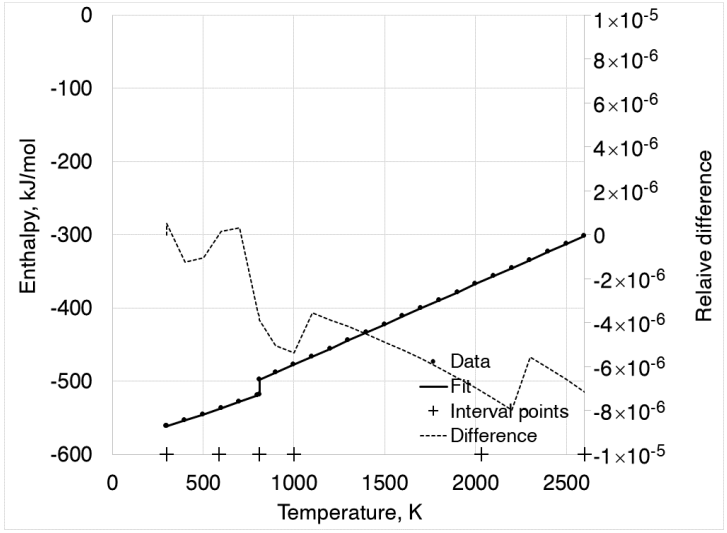


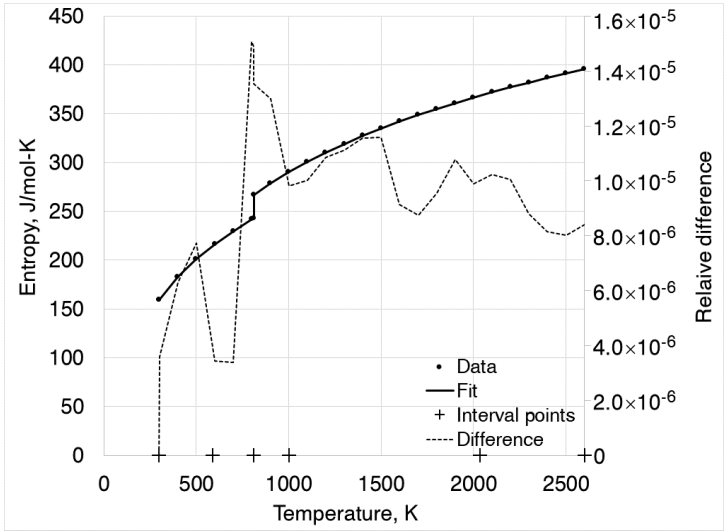


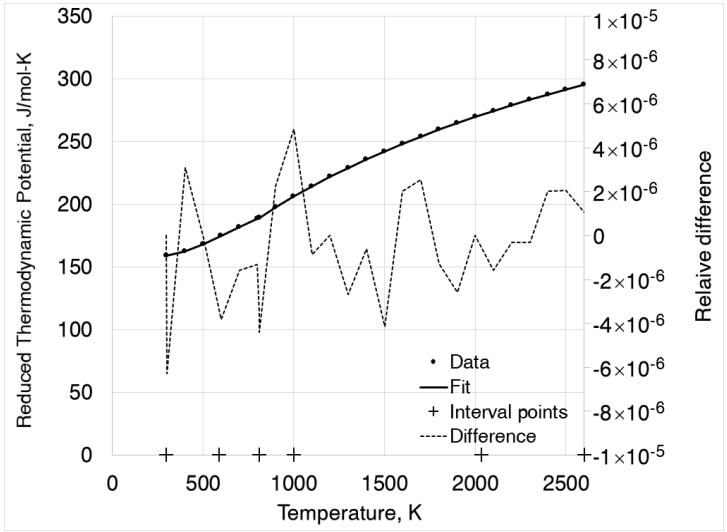
Strontium Iodide (SrI_2) – I2Sr1(cr,l)

MW = 341.43 g/mol – $\Delta_f H^o$(298.15 K) = -561.493 kJ/mol – S^o(298.15 K) = 159.12 J/mol-K – p^o = 0.1 MPa



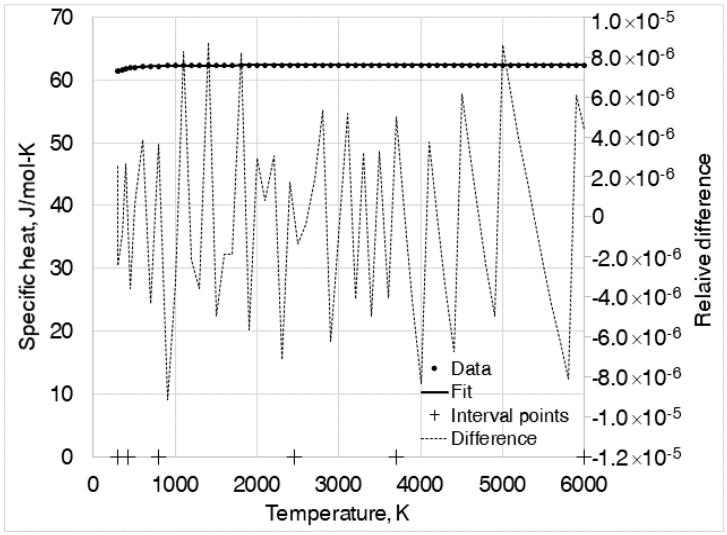


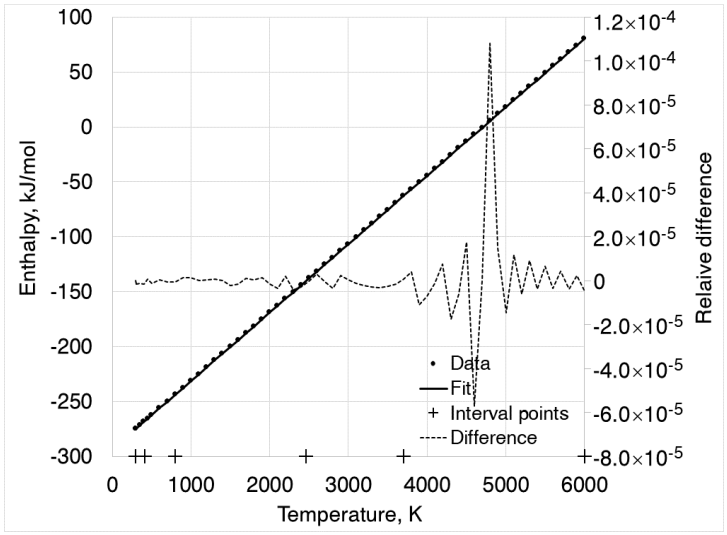


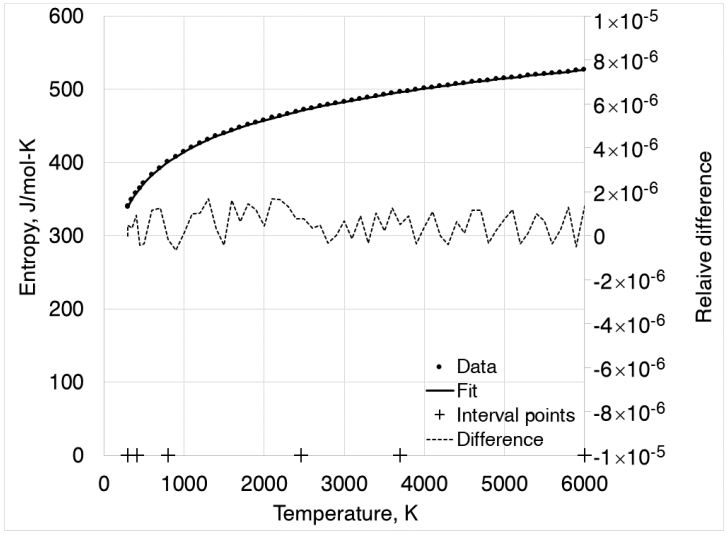


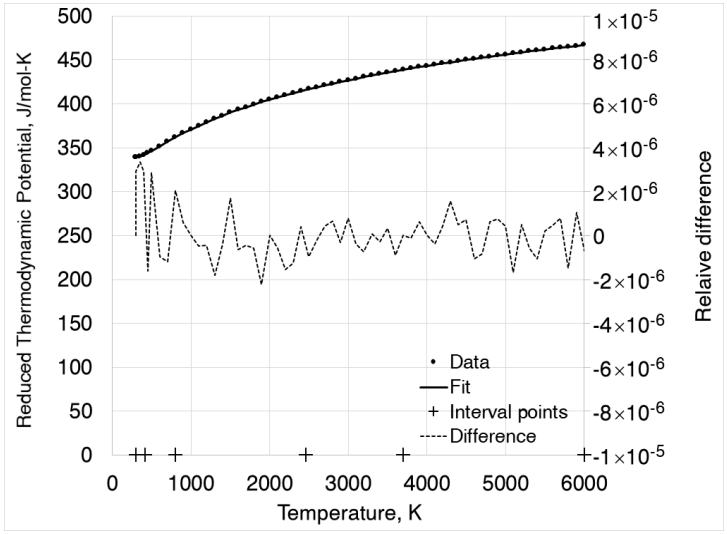
Strontium Iodide (SrI_2) – I2Sr1(g)

MW = 341.43 g/mol – $\Delta_f H^o$(298.15 K) = -274.889 kJ/mol – S^o(298.15 K) = 339.554 J/mol-K – p^o = 0.1 MPa



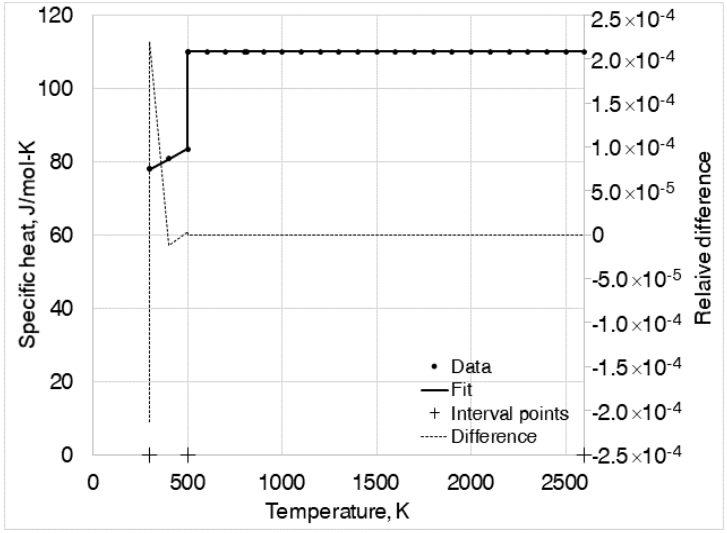


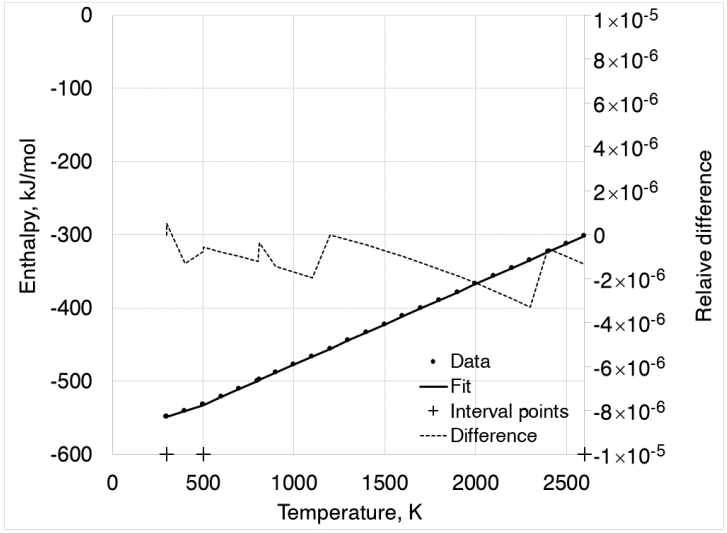


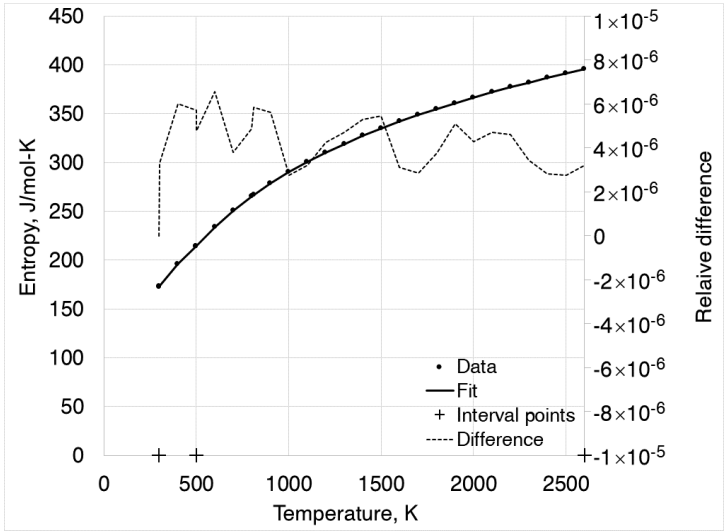


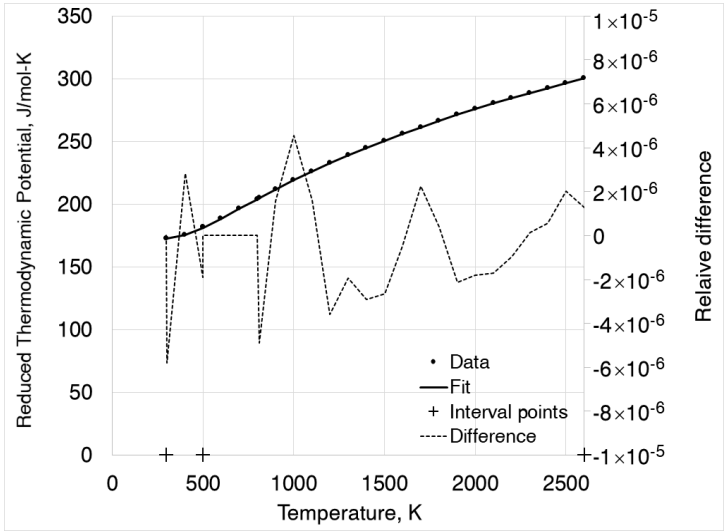
Strontium Iodide (SrI_2) – I2Sr1(l)

MW = 341.43 g/mol – $\Delta_f H^o$(298.15 K) = -548.748 kJ/mol – S^o(298.15 K) = 172.44 J/mol-K – p^o = 0.1 MPa



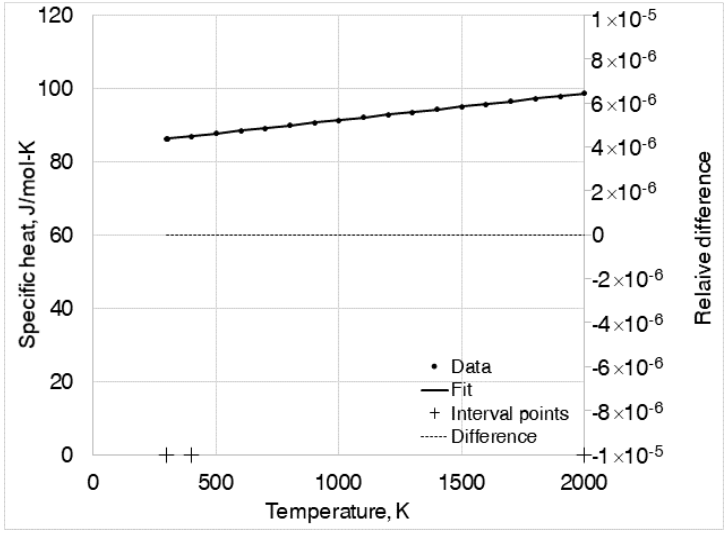


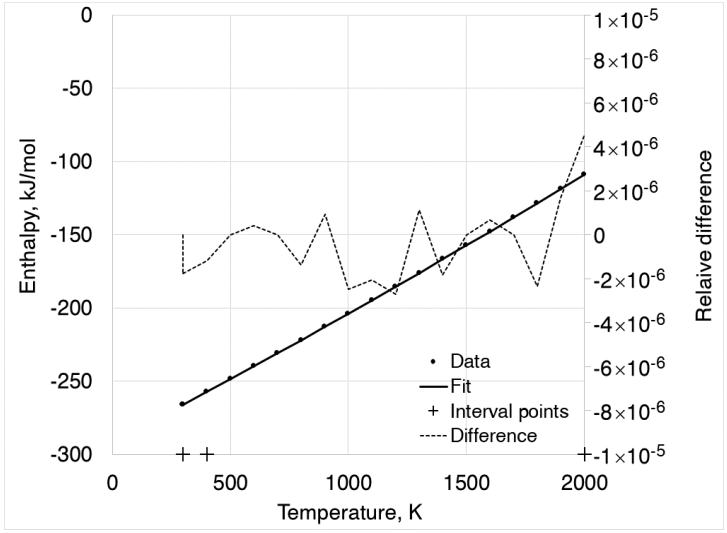


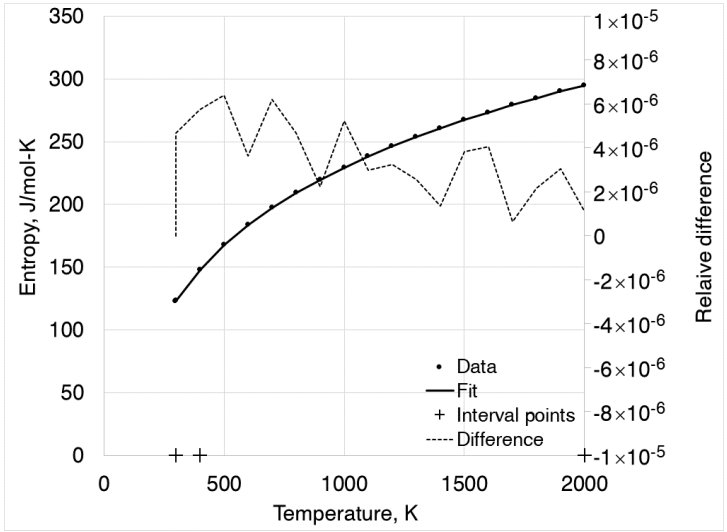


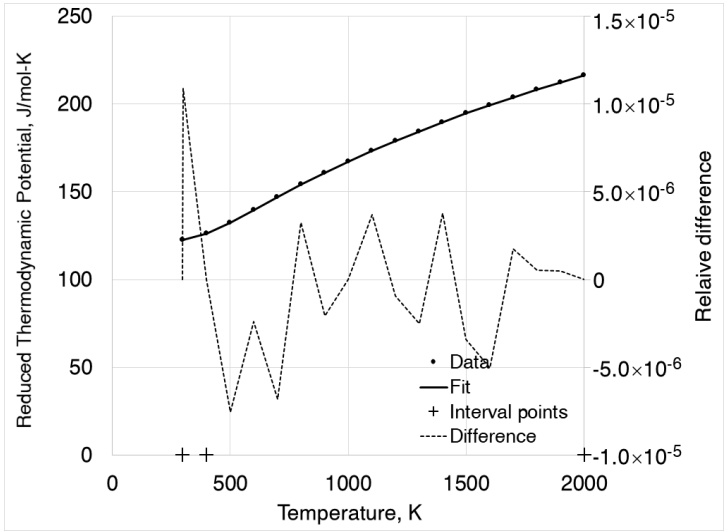
Titanium Iodide (TiI_2) – I2Ti1(cr)

MW = 301.69 g/mol – $\Delta_f H^o$(298.15 K) = -266.102 kJ/mol – S^o(298.15 K) = 122.591 J/mol-K – p^o = 0.1 MPa



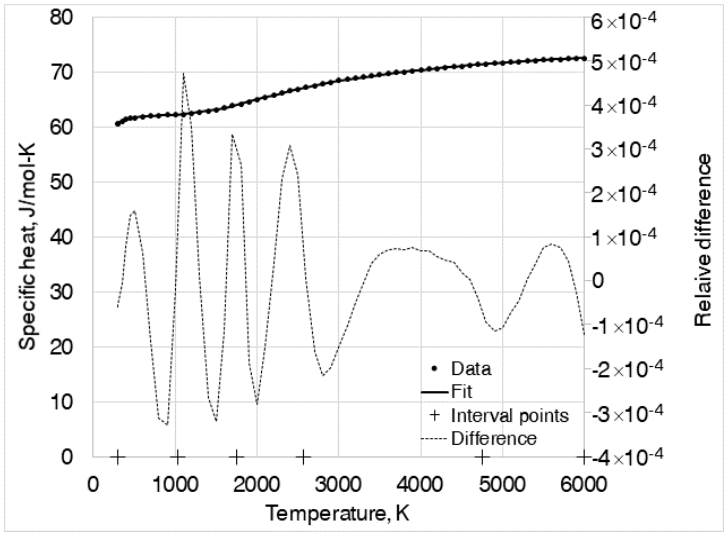


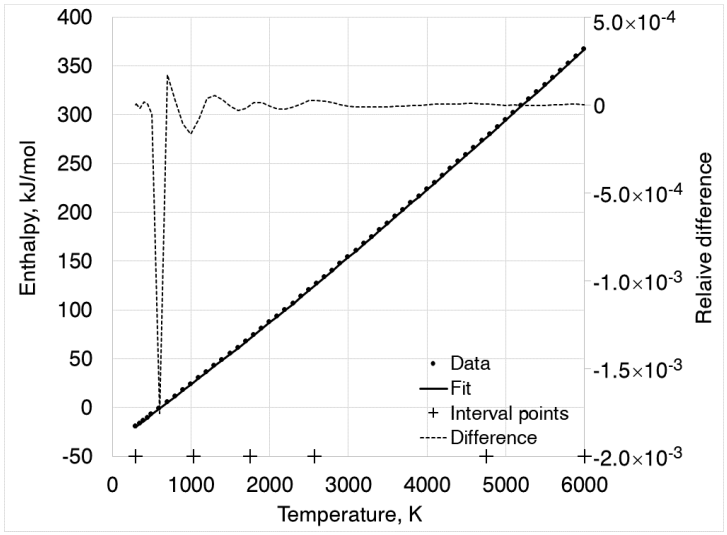


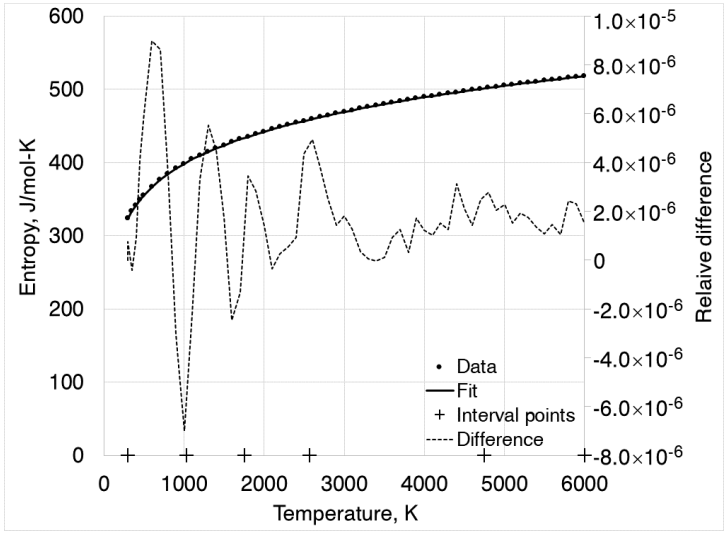


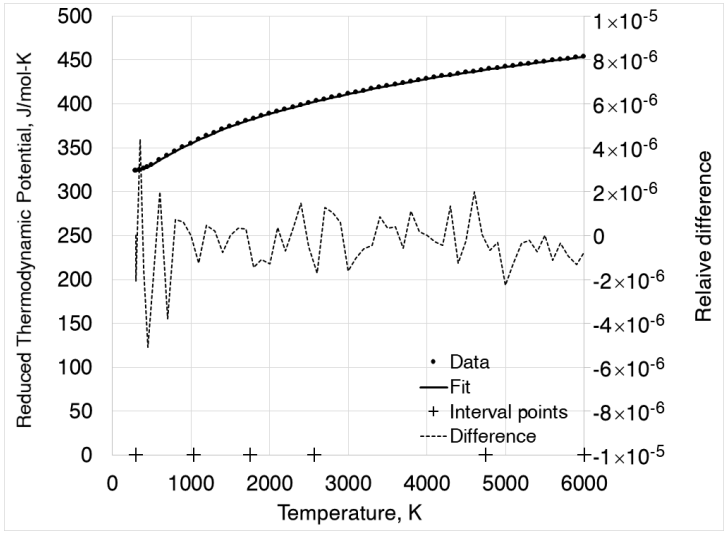
Titanium Iodide (TiI_2) – I2Ti1(g)

MW = 301.69 g/mol – $\Delta_f H^o$(298.15 K) = -19.665 kJ/mol – S^o(298.15 K) = 323.673 J/mol-K – p^o = 0.1 MPa



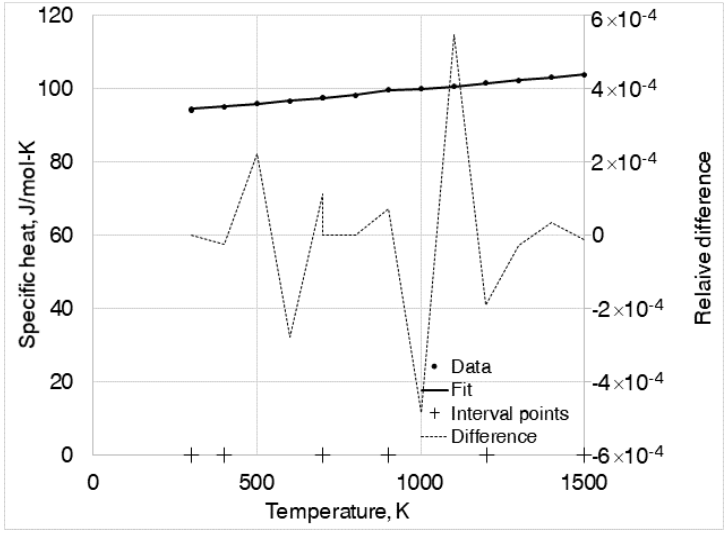


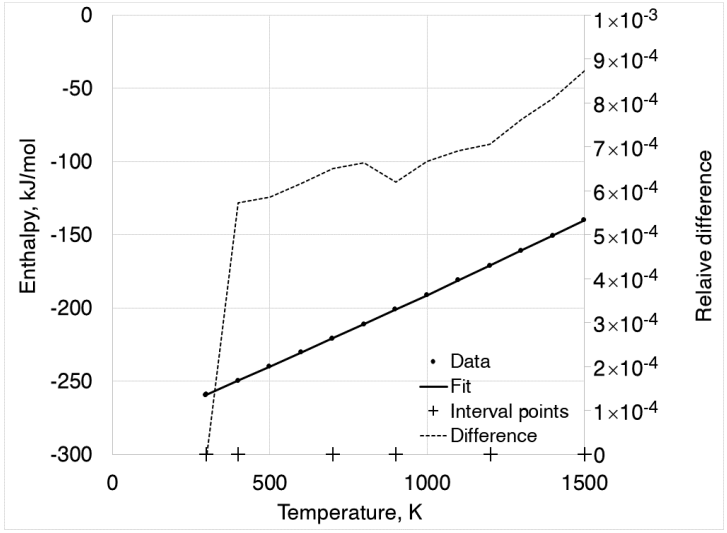


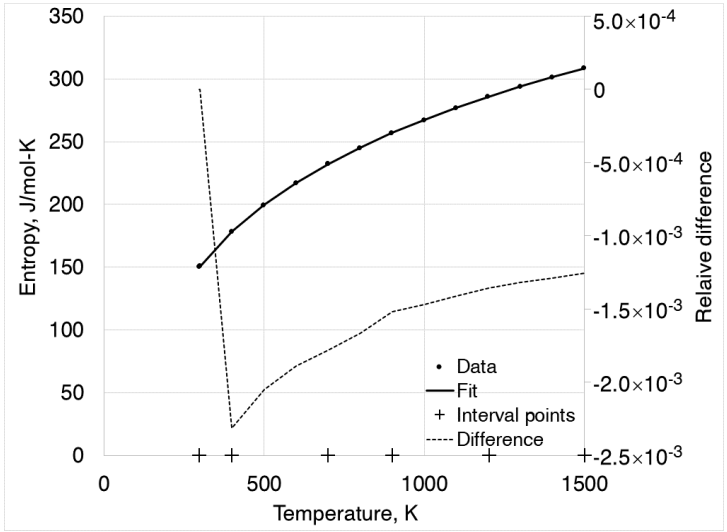


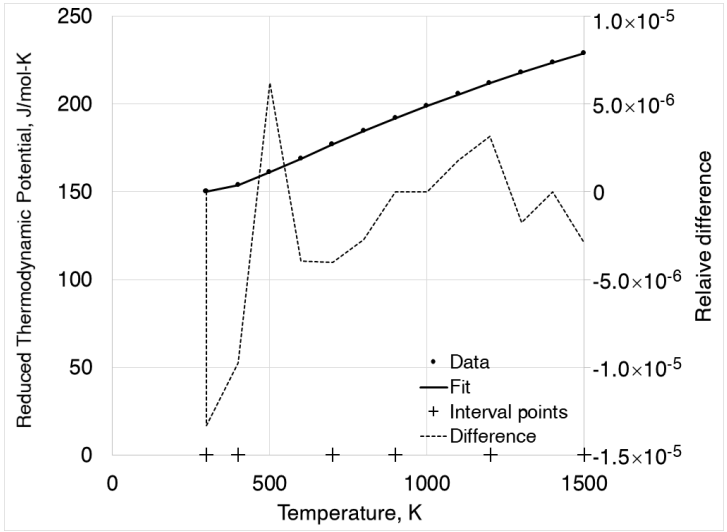
Zirconium Iodide (ZrI_2) – I2Zr1(cr)

MW = 345.03 g/mol – $\Delta_f H^o$(298.15 K) = -259.408 kJ/mol – S^o(298.15 K) = 150.206 J/mol-K – p^o = 0.1 MPa



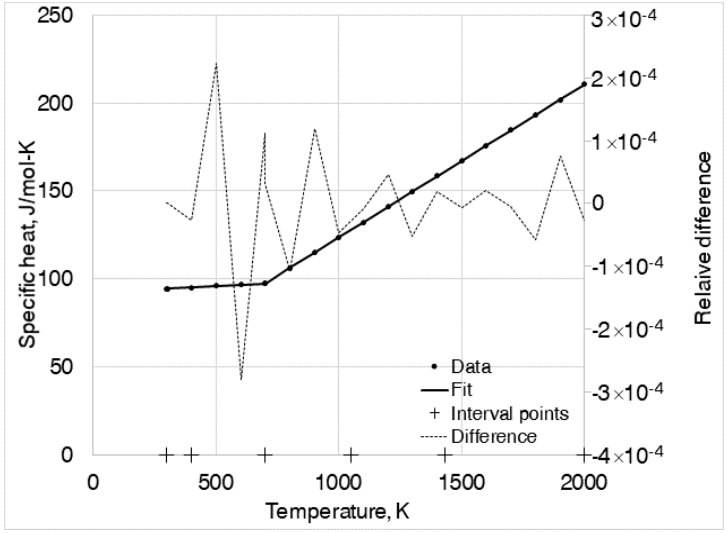


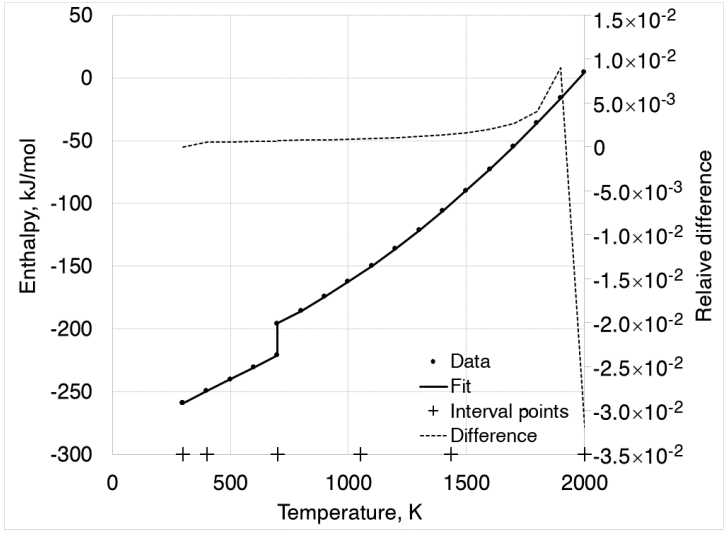


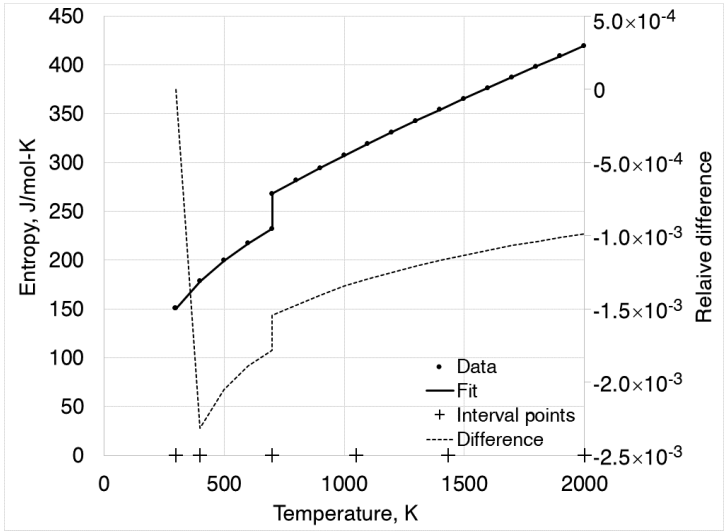


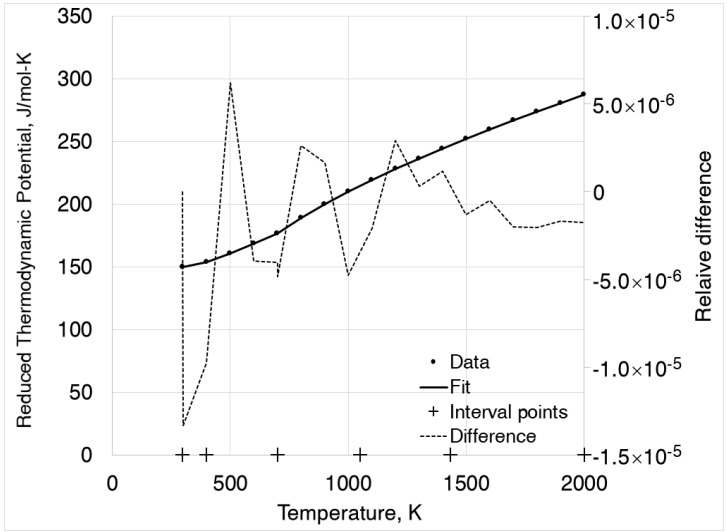
Zirconium Iodide (ZrI_2) – I2Zr1(cr,l)

MW = 345.03 g/mol – $\Delta_f H^o$(298.15 K) = -259.408 kJ/mol – S^o(298.15 K) = 150.206 J/mol-K – p^o = 0.1 MPa



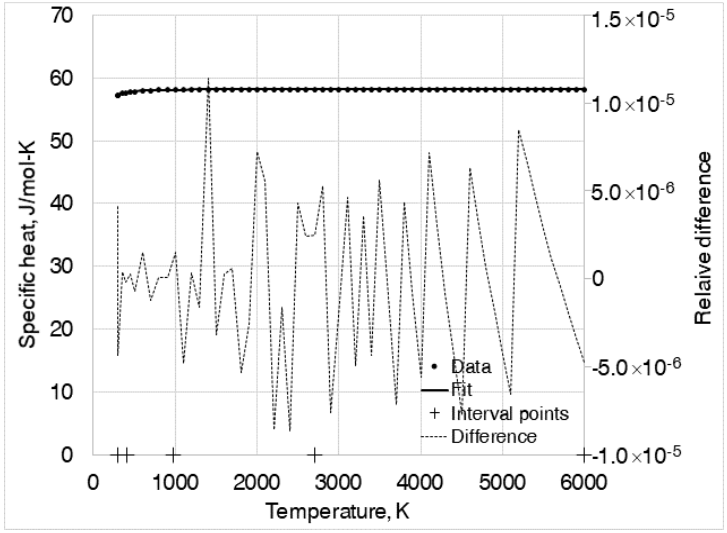


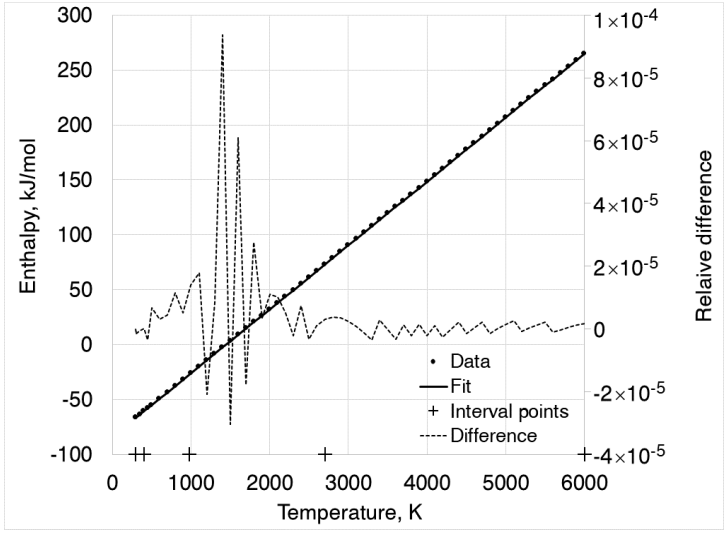


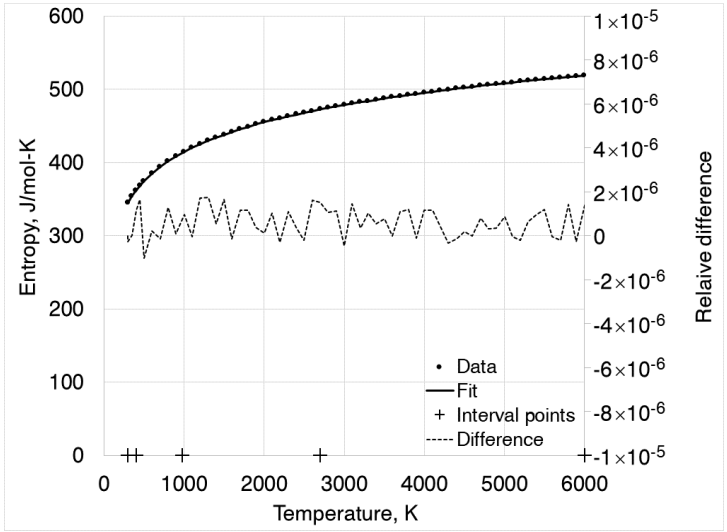


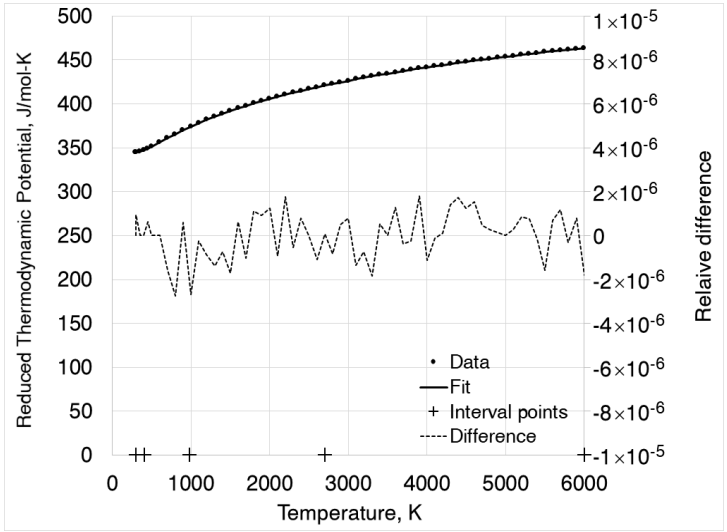
Zirconium Iodide (ZrI_2) – I2Zr1(g)

MW = 345.03 g/mol – $\Delta_f H^o$(298.15 K) = -66.643 kJ/mol – S^o(298.15 K) = 344.782 J/mol-K – p^o = 0.1 MPa



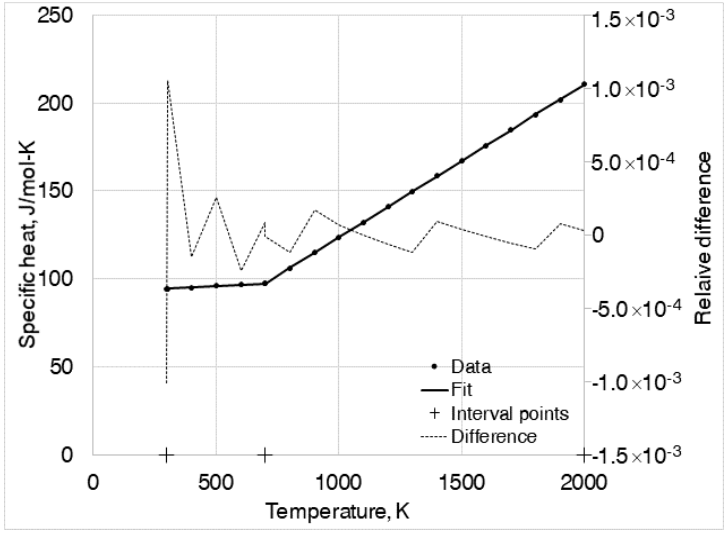


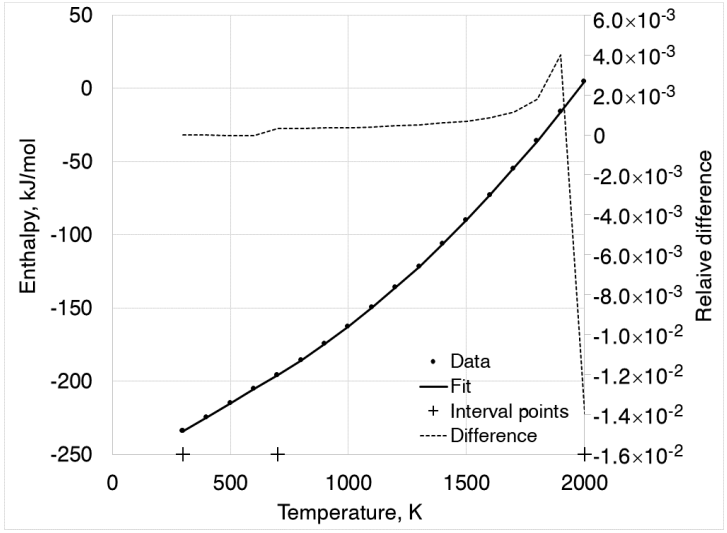


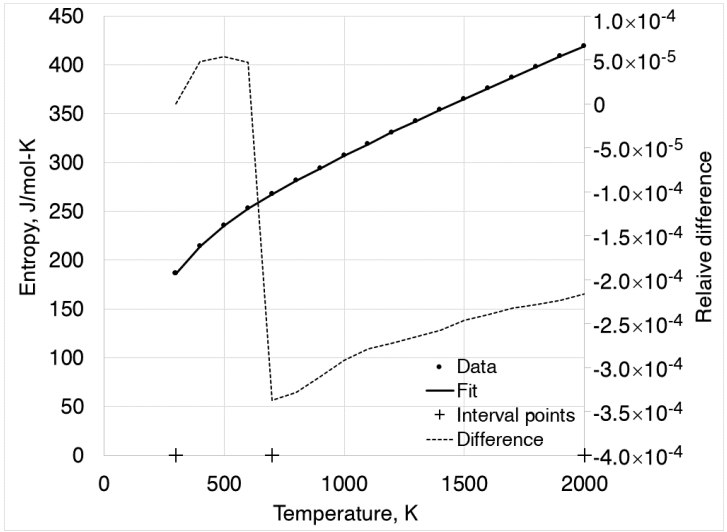


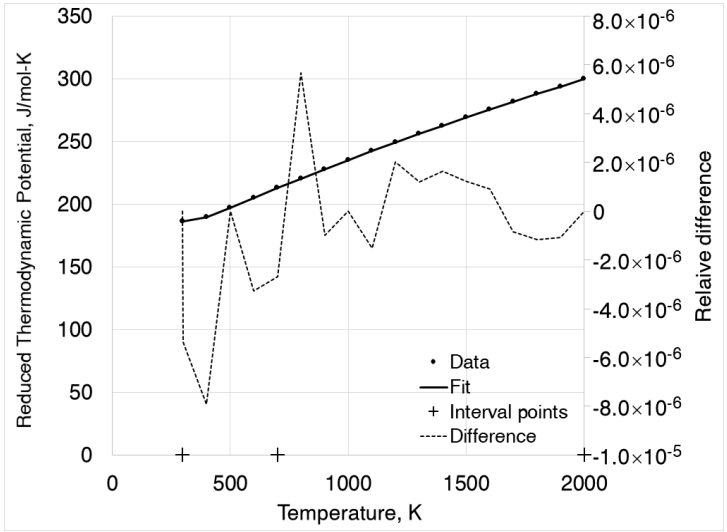
Zirconium Iodide (ZrI_2) – I2Zr1(l)

MW = 345.03 g/mol – $\Delta_f H^o$(298.15 K) = -234.239 kJ/mol – S^o(298.15 K) = 186.168 J/mol-K – p^o = 0.1 MPa



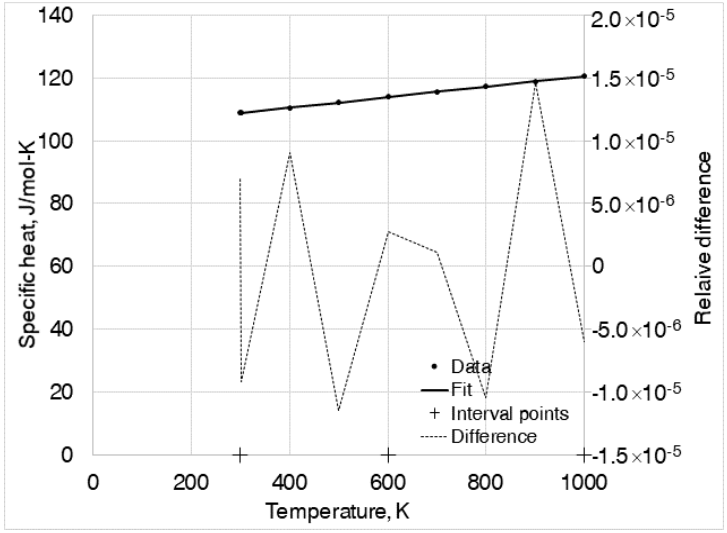


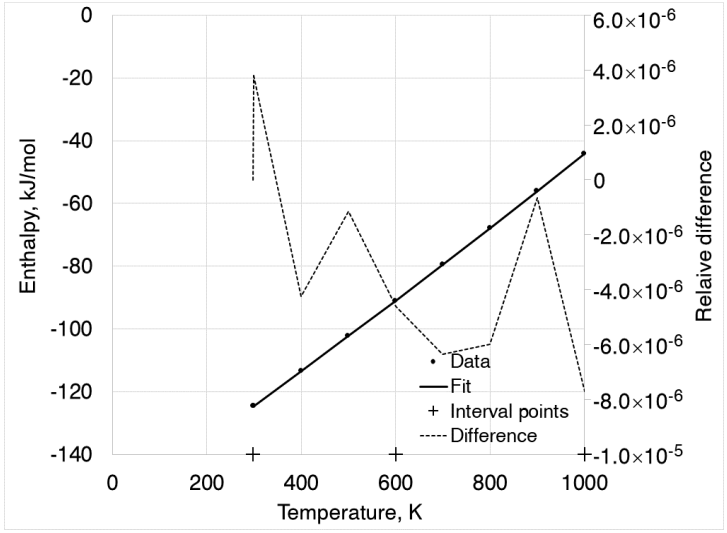


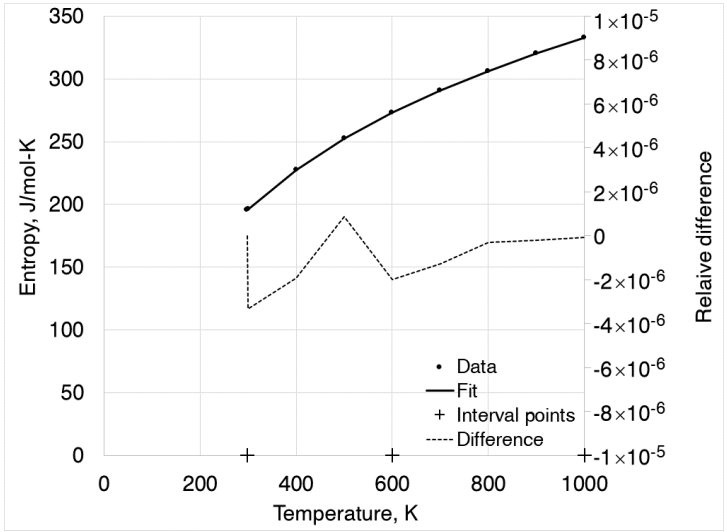


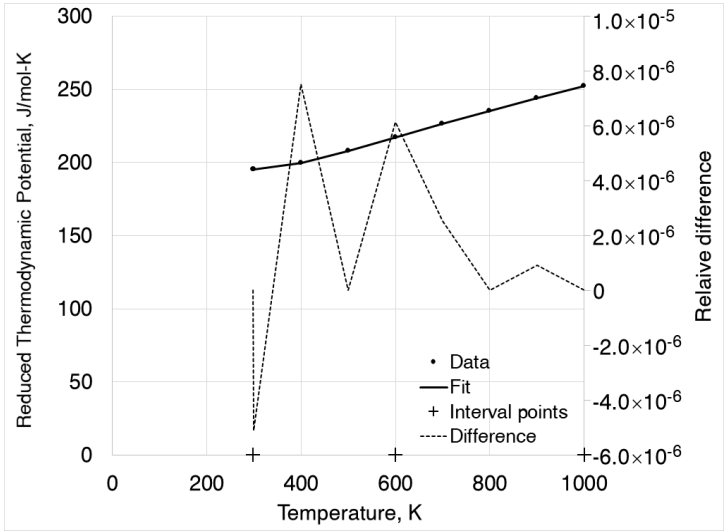
Molybdenum Iodide (MoI_3) – I3Mo1(cr)

MW = 476.65 g/mol – $\Delta_f H^o$(298.15 K) = -124.683 kJ/mol – S^o(298.15 K) = 195.393 J/mol-K – p^o = 0.1 MPa



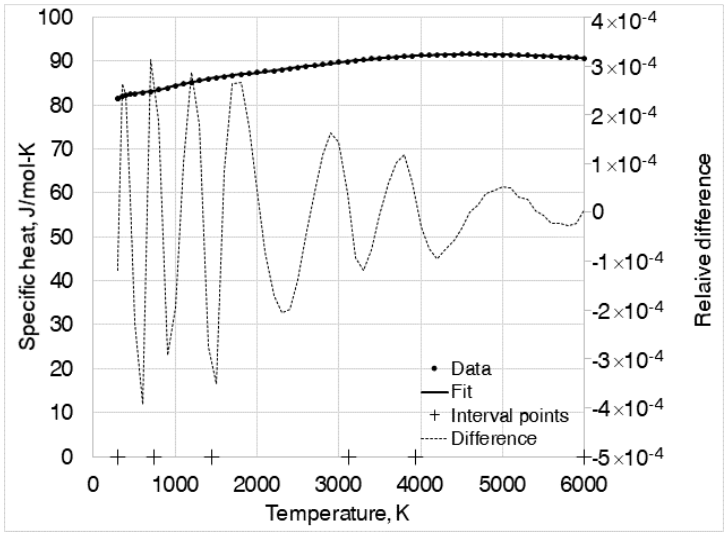


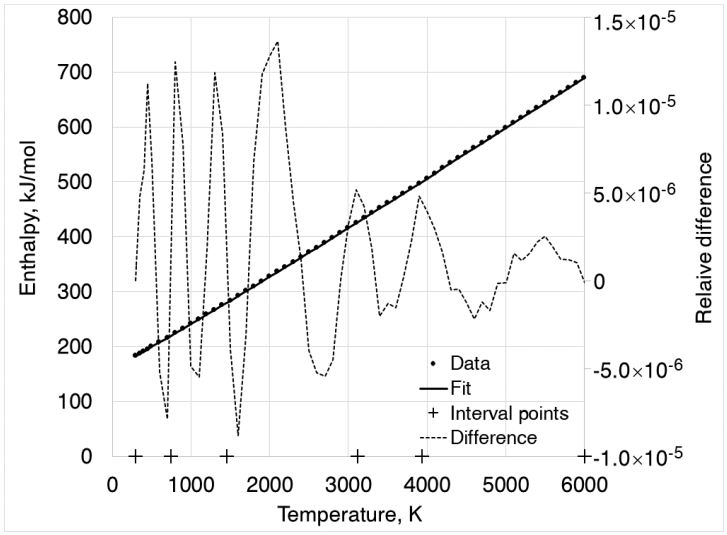


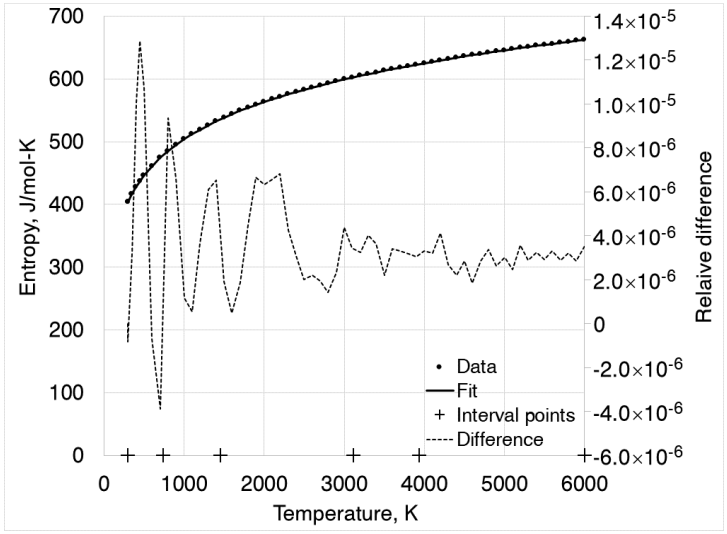


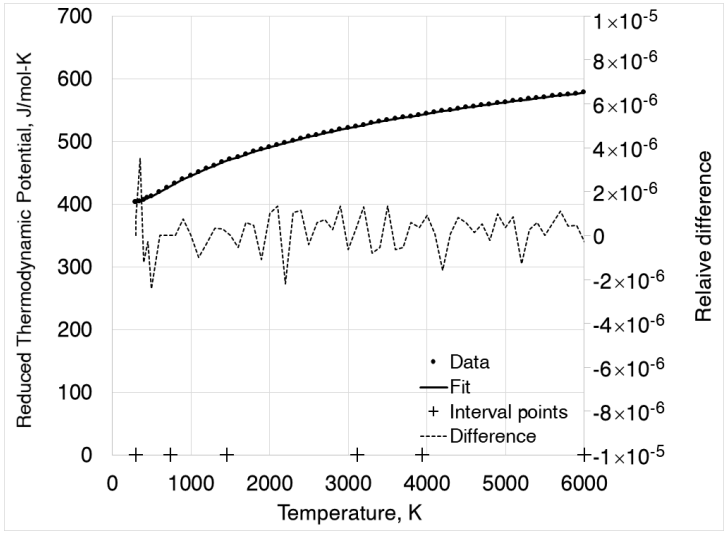
Molybdenum Iodide (MoI_3) – I3Mo1(g)

MW = 476.65 g/mol – $\Delta_f H^o$(298.15 K) = 182.841 kJ/mol – S^o(298.15 K) = 403.532 J/mol-K – p^o = 0.1 MPa



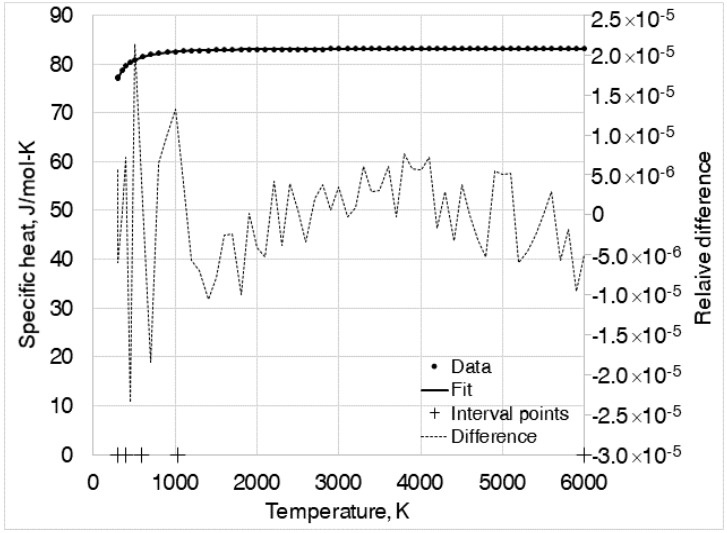


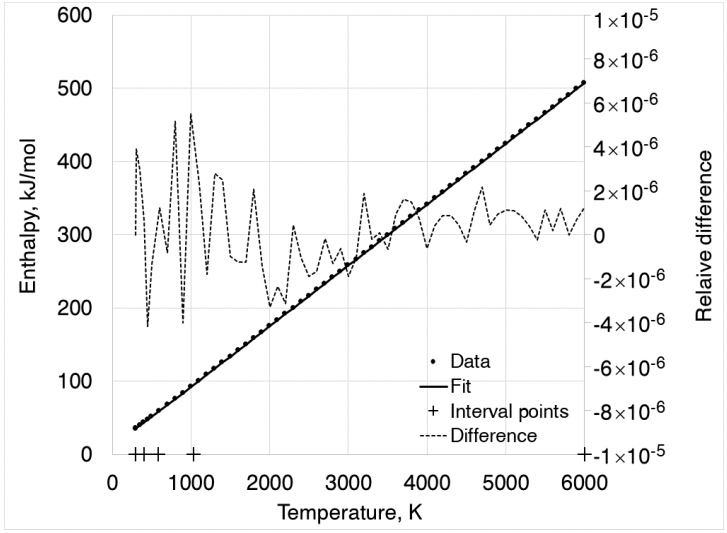


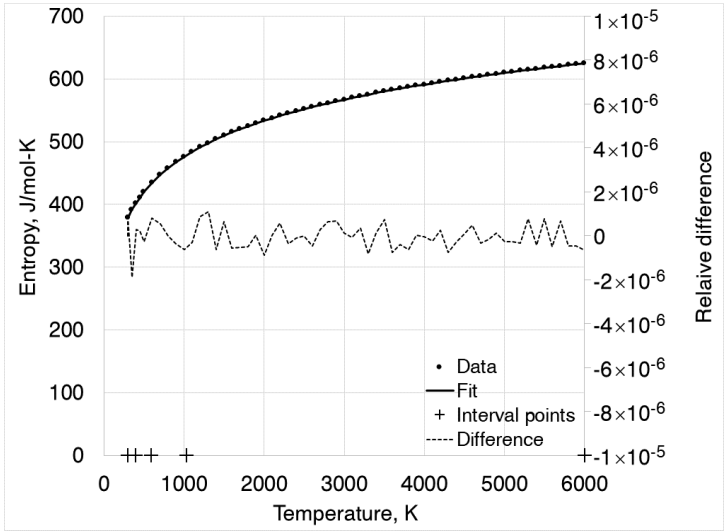


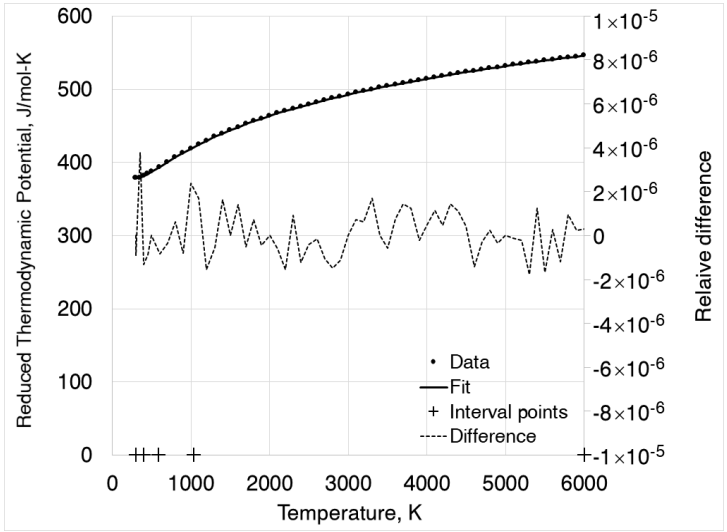
Triiodosilyl (SiI_3) – I3Si1(g)

MW = 408.799 g/mol – $\Delta_f H^o$(298.15 K) = 35.314 kJ/mol – S^o(298.15 K) = 378.305 J/mol-K – p^o = 0.1 MPa



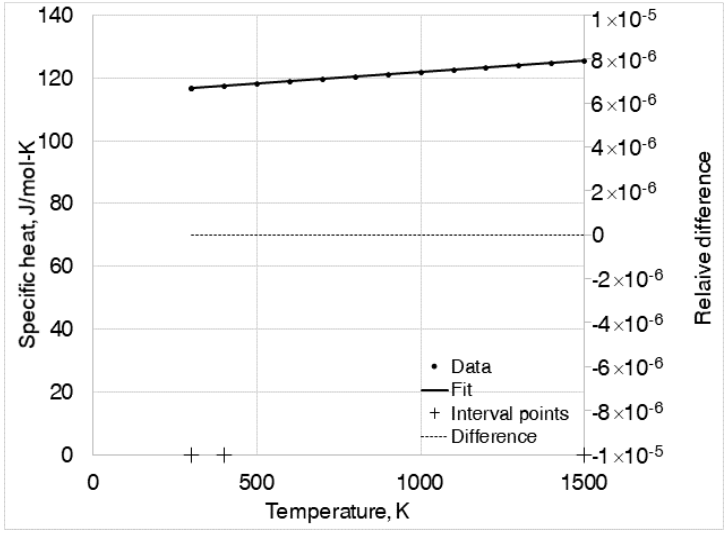


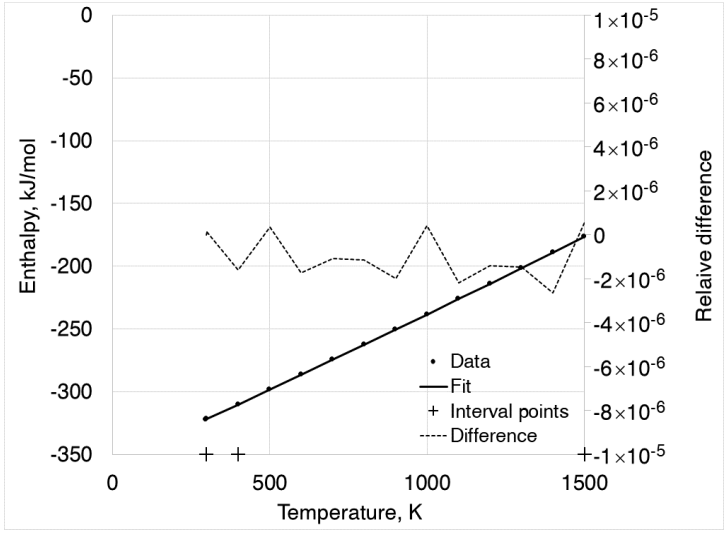


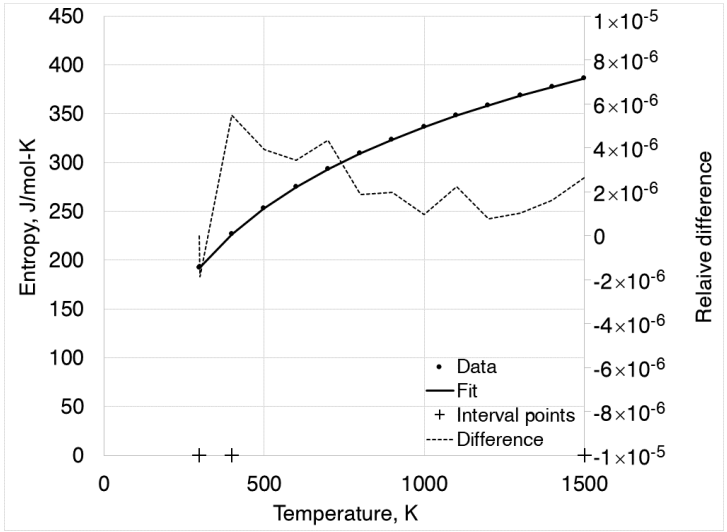


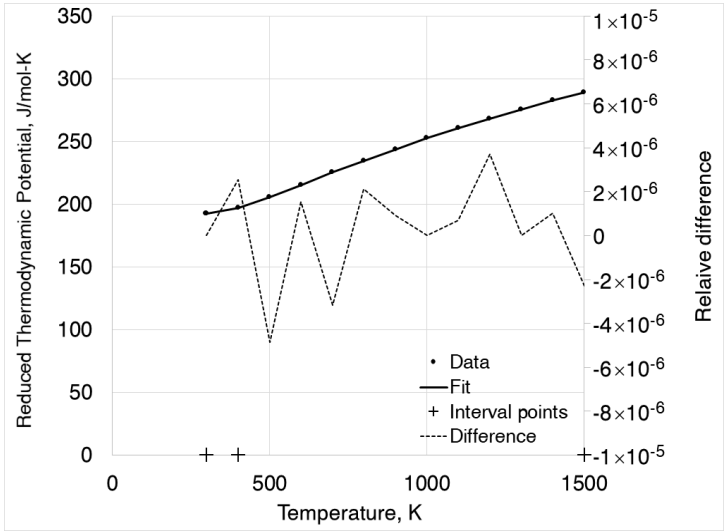
Titanium Iodide (TiI_3) – I3Ti1(cr)

MW = 428.59 g/mol – $\Delta_f H^o$(298.15 K) = -322.168 kJ/mol – S^o(298.15 K) = 192.464 J/mol-K – p^o = 0.1 MPa



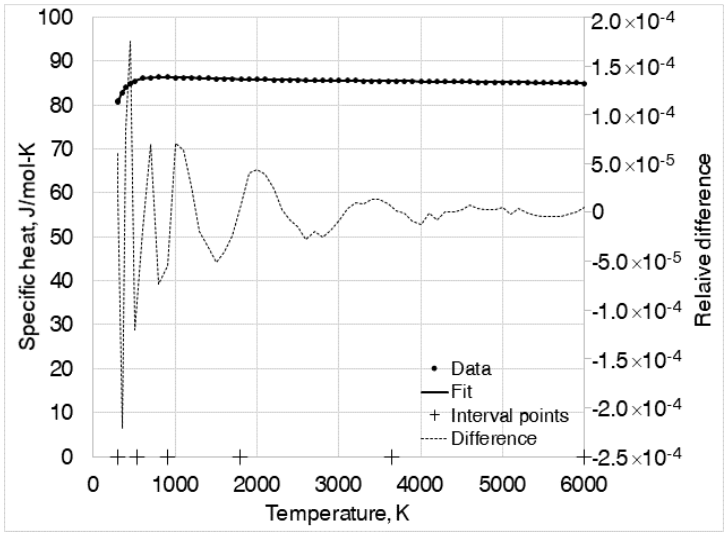


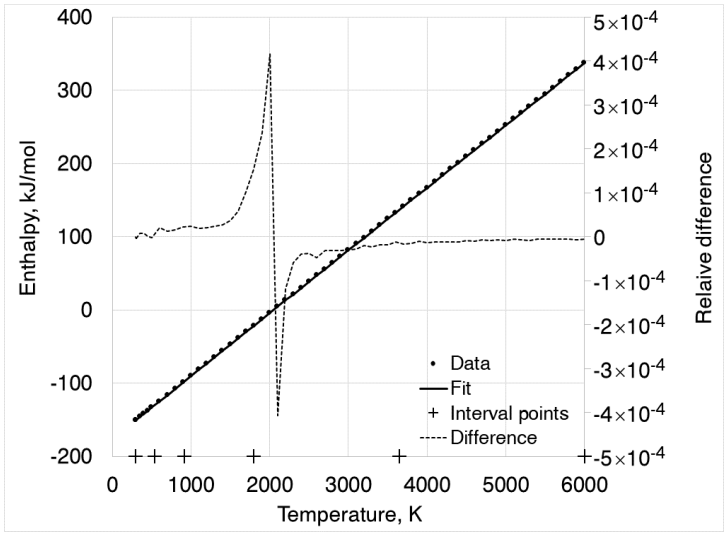


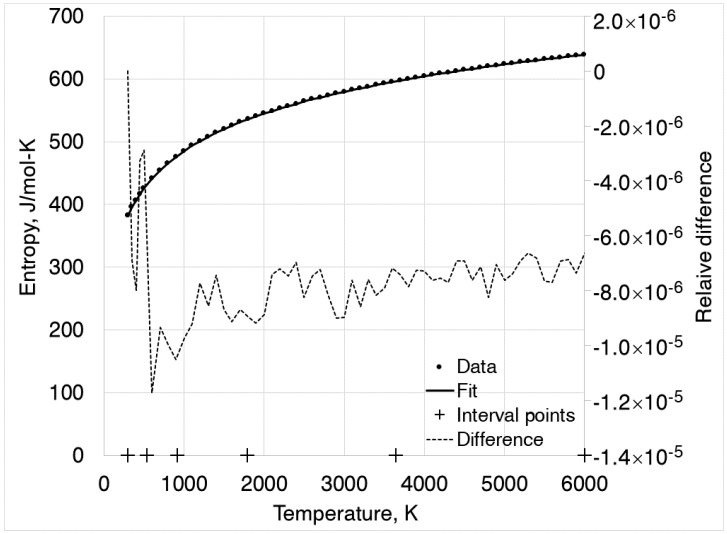


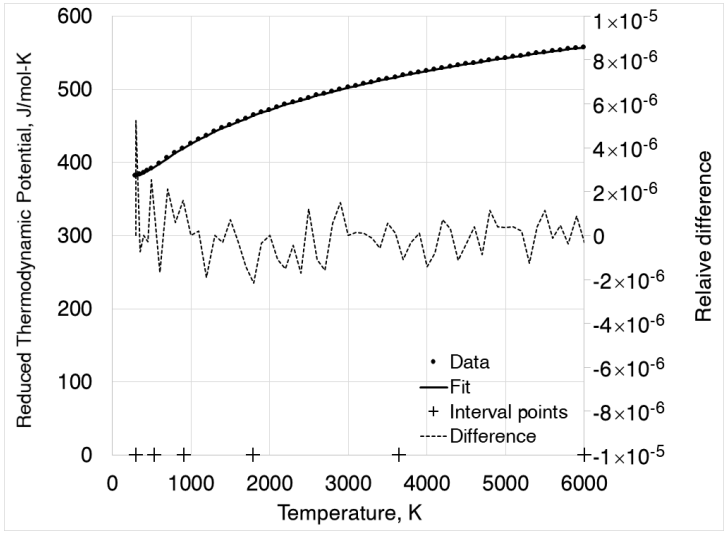
Titanium Iodide (TiI_3) – I3Ti1(g)

MW = 428.59 g/mol – $\Delta_f H^o$(298.15 K) = -150.206 kJ/mol – S^o(298.15 K) = 382.164 J/mol-K – p^o = 0.1 MPa



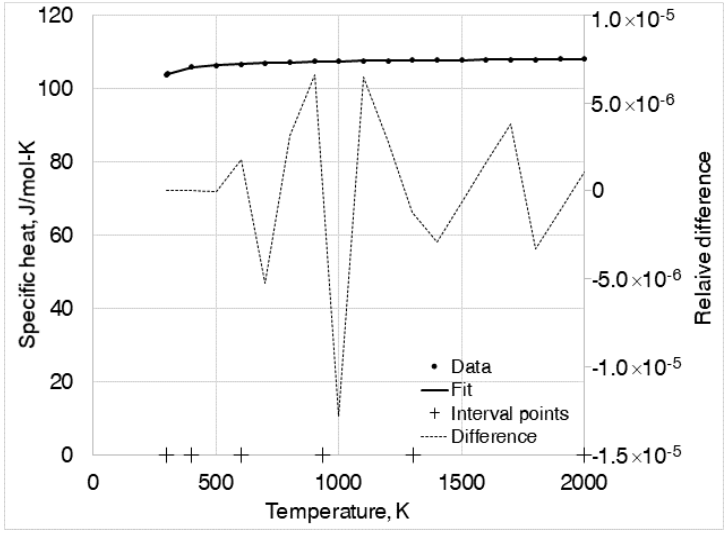


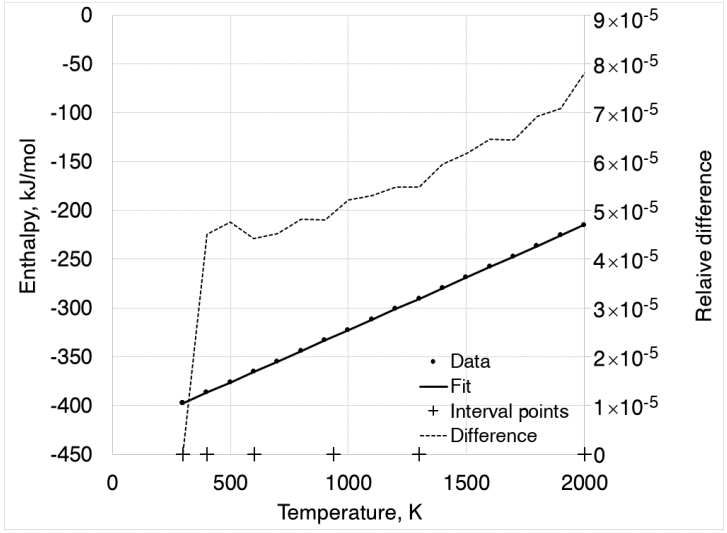


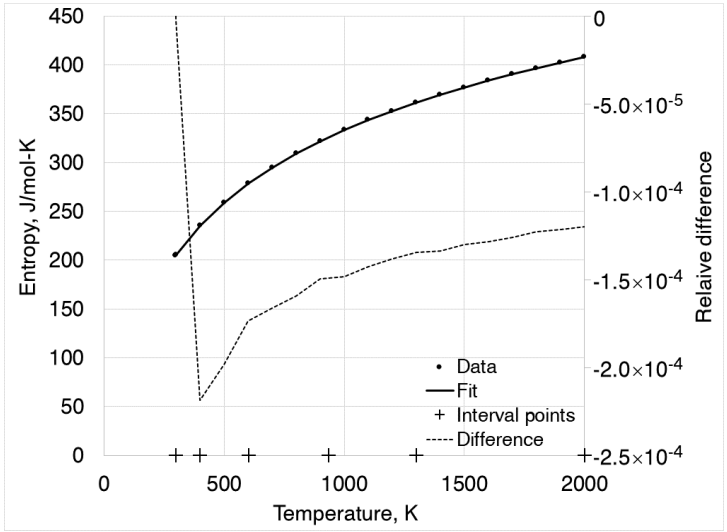


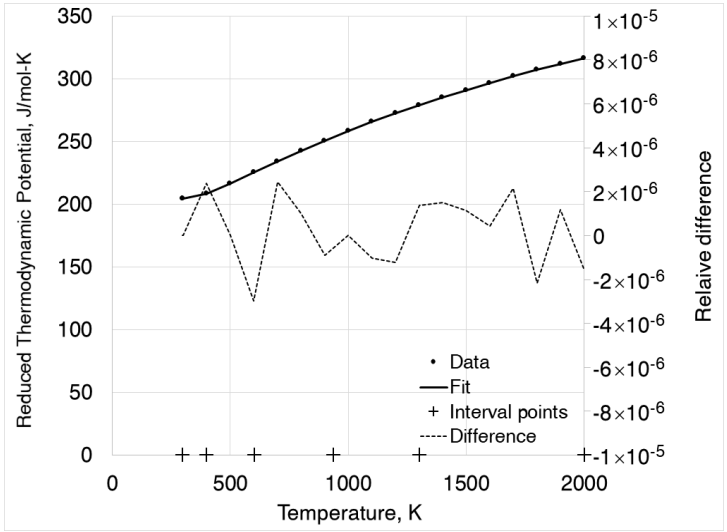
Zirconium Iodide (ZrI_3) – I3Zr1(cr)

MW = 471.93 g/mol – $\Delta_f H^o$(298.15 K) = -397.48 kJ/mol – S^o(298.15 K) = 204.585 J/mol-K – p^o = 0.1 MPa



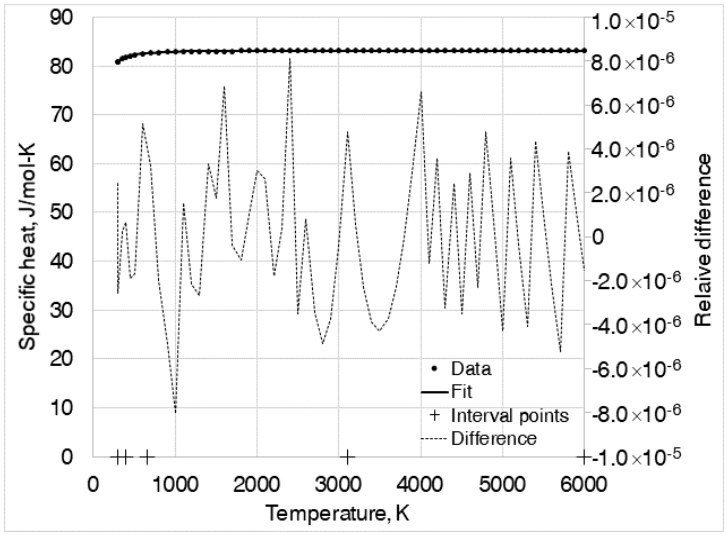


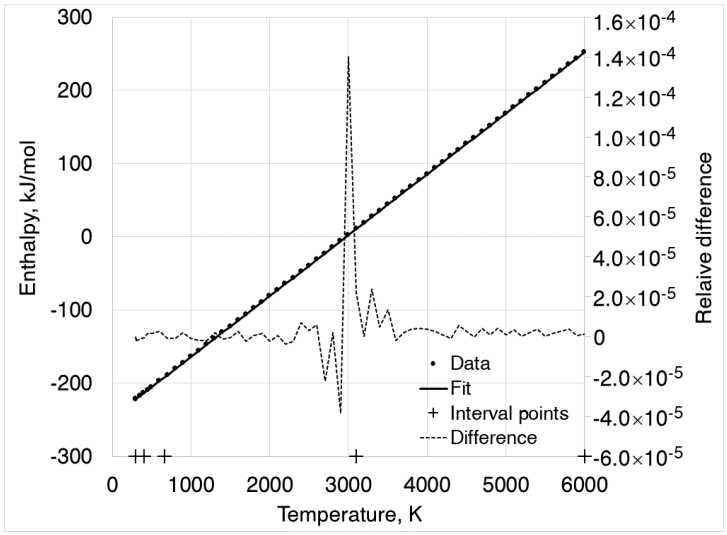


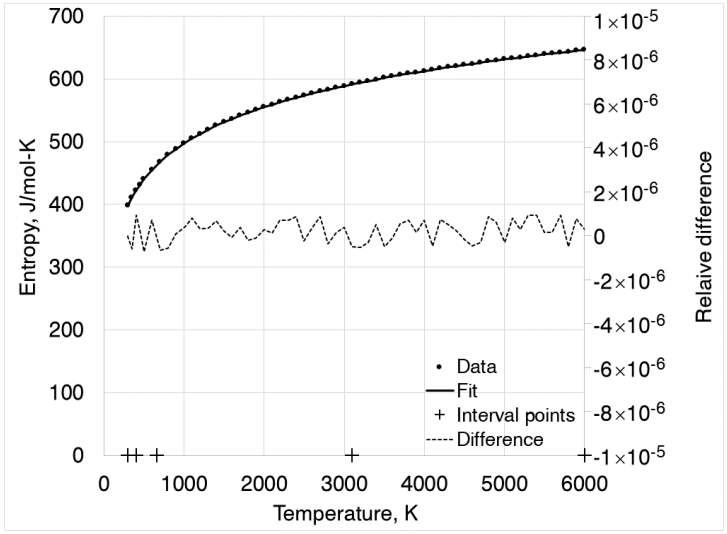


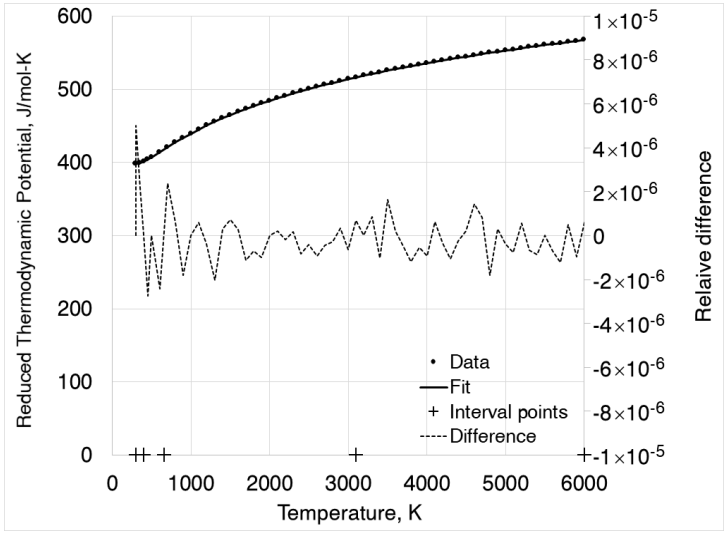
Zirconium Iodide (ZrI_3) – I3Zr1(g)

MW = 471.93 g/mol – $\Delta_f H^o$(298.15 K) = -221.752 kJ/mol – S^o(298.15 K) = 397.776 J/mol-K – p^o = 0.1 MPa



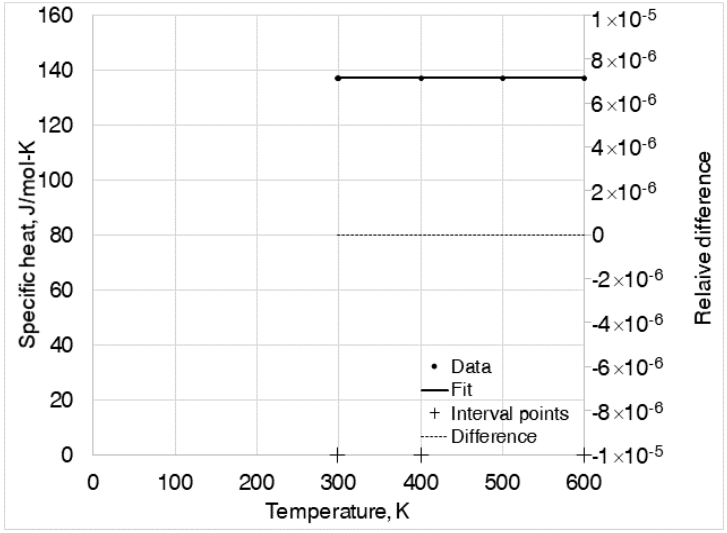


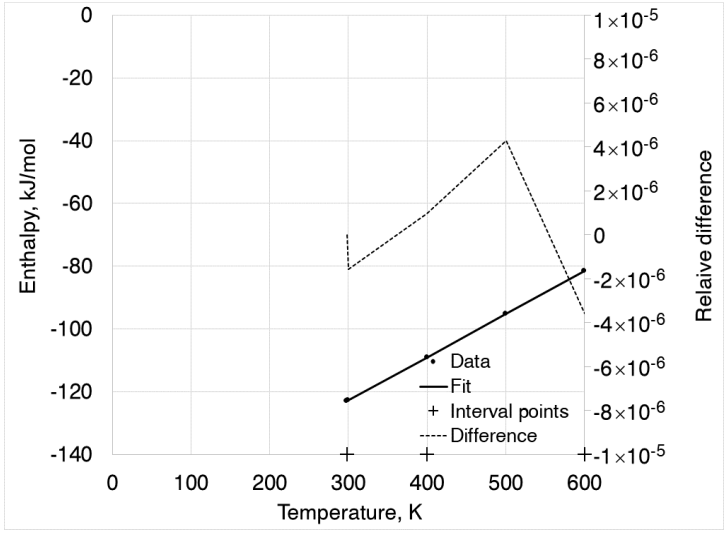


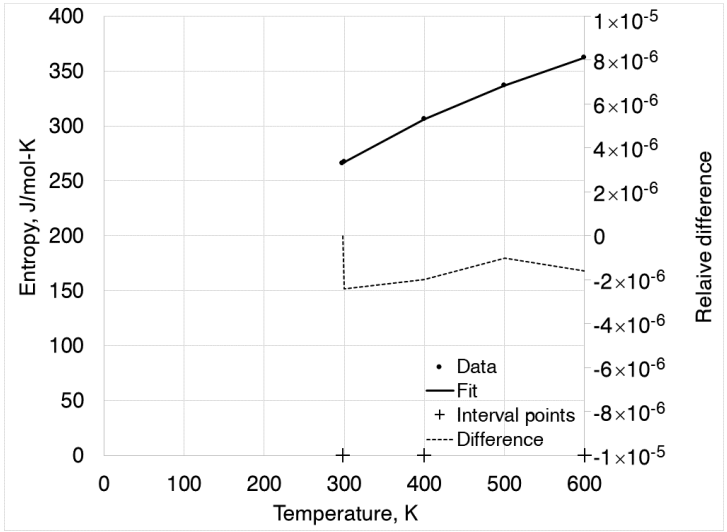


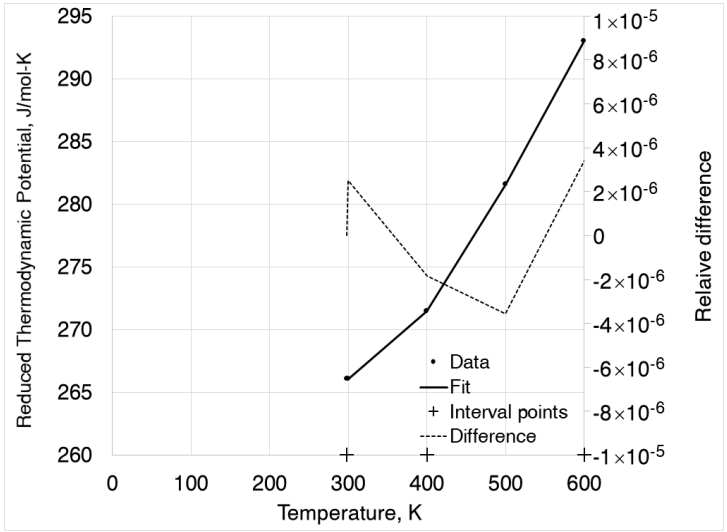
Molybdenum Iodide (MoI_4) – I4Mo1(cr)

MW = 603.56 g/mol – $\Delta_f H^o$(298.15 K) = -123.01 kJ/mol – S^o(298.15 K) = 266.061 J/mol-K – p^o = 0.1 MPa



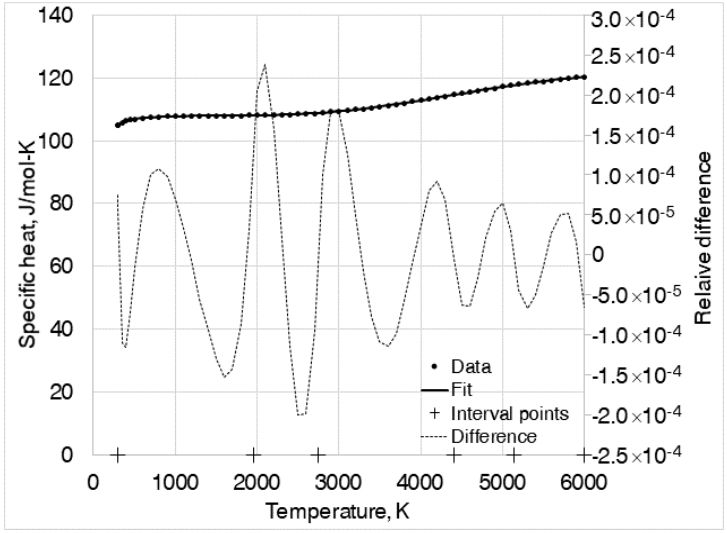


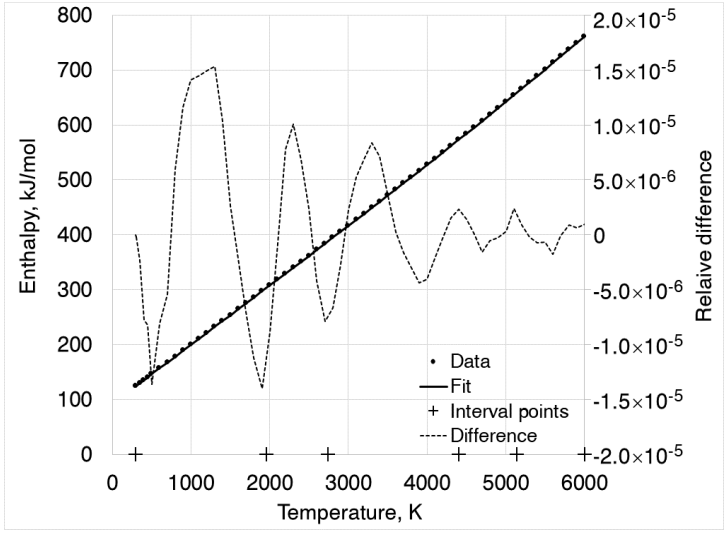


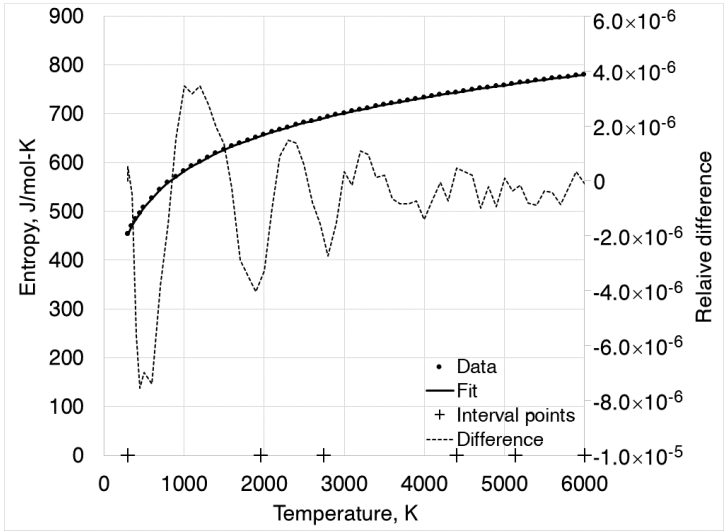


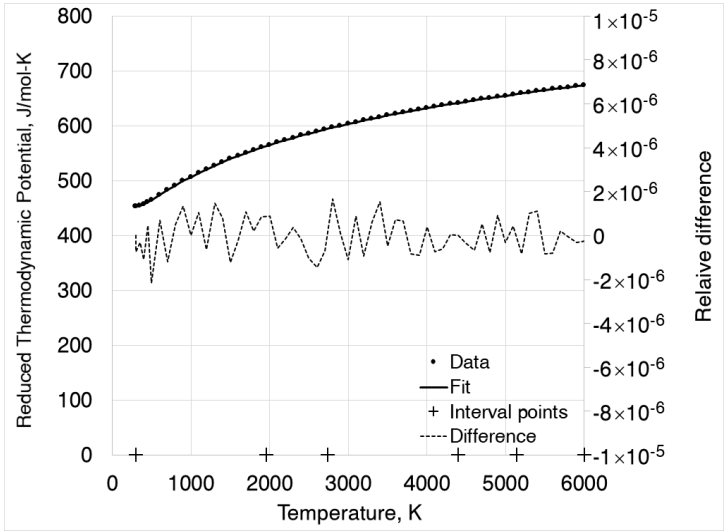
Molybdenum Iodide (MoI_4) – I4Mo1(g)

MW = 603.56 g/mol – $\Delta_f H^o$(298.15 K) = 124.683 kJ/mol – S^o(298.15 K) = 452.812 J/mol-K – p^o = 0.1 MPa



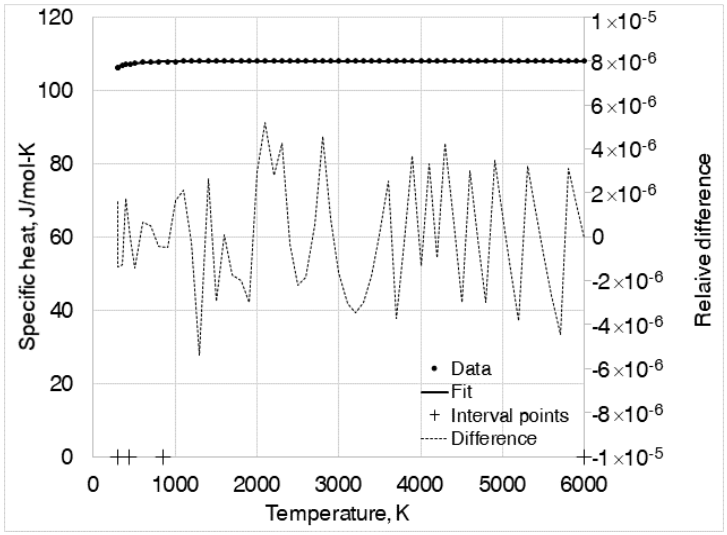


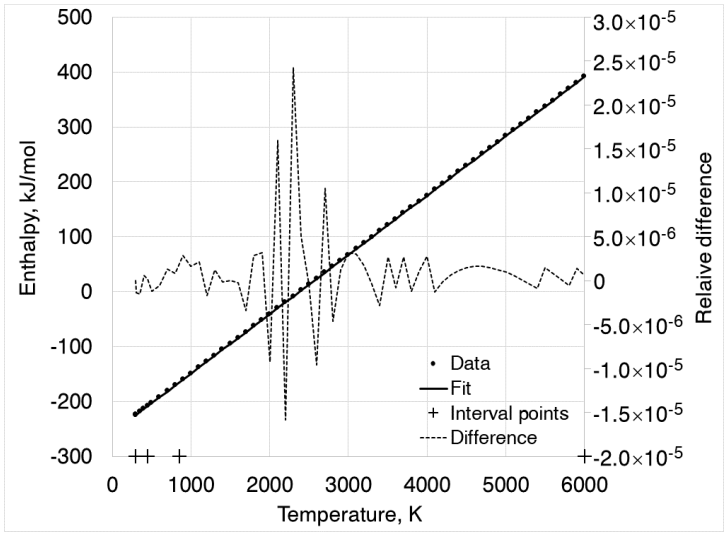


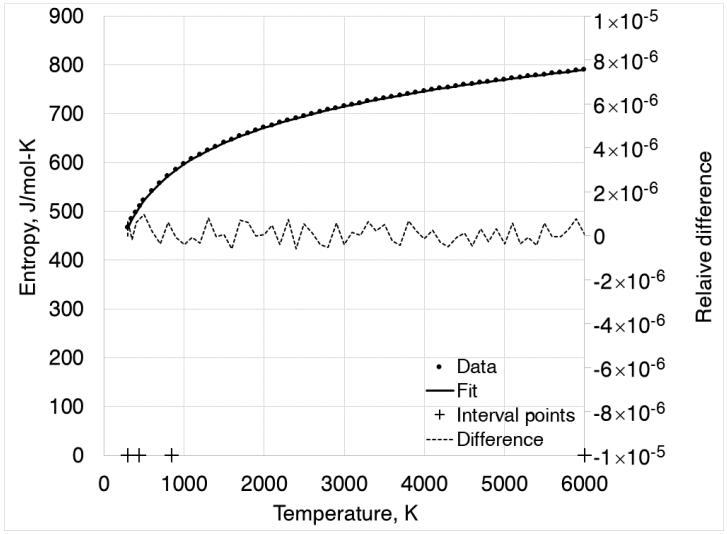


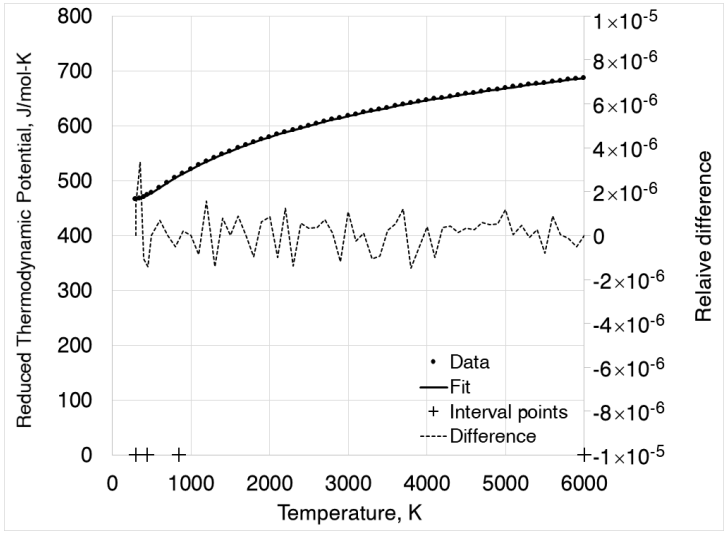
Lead Iodide (PbI_4) – I4Pb1(g)

MW = 714.8 g/mol – $\Delta_f H^o$(298.15 K) = -224.467 kJ/mol – S^o(298.15 K) = 466.266 J/mol-K – p^o = 0.1 MPa



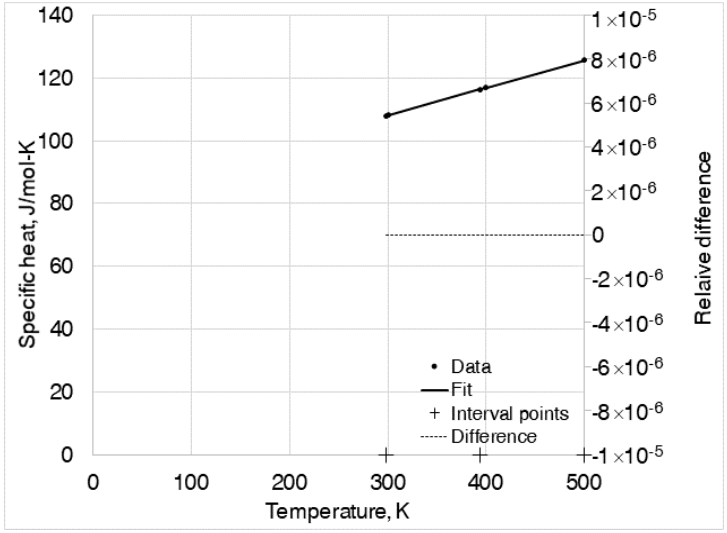


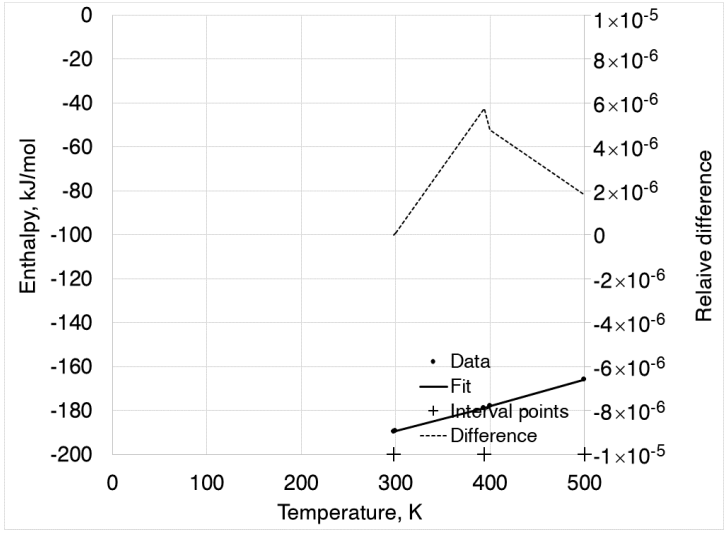


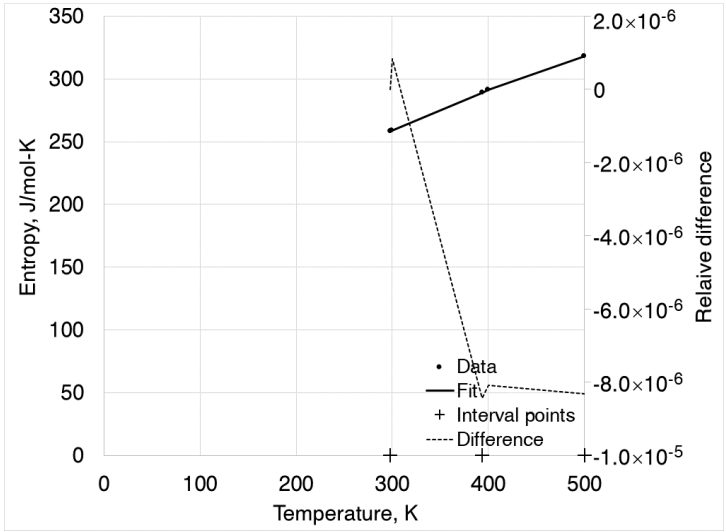


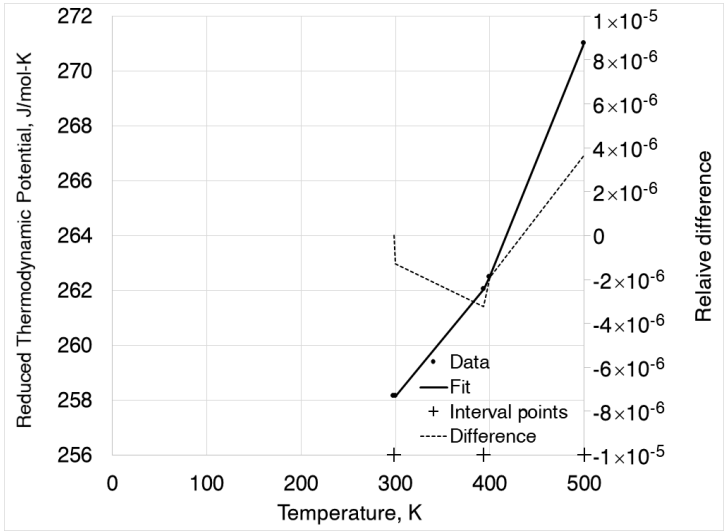
Tetraiodosilane (SiI_4) – I4Si1(cr)

MW = 535.7035 g/mol – $\Delta_f H^o$(298.15 K) = -189.535 kJ/mol – S^o(298.15 K) = 258.153 J/mol-K – p^o = 0.1 MPa



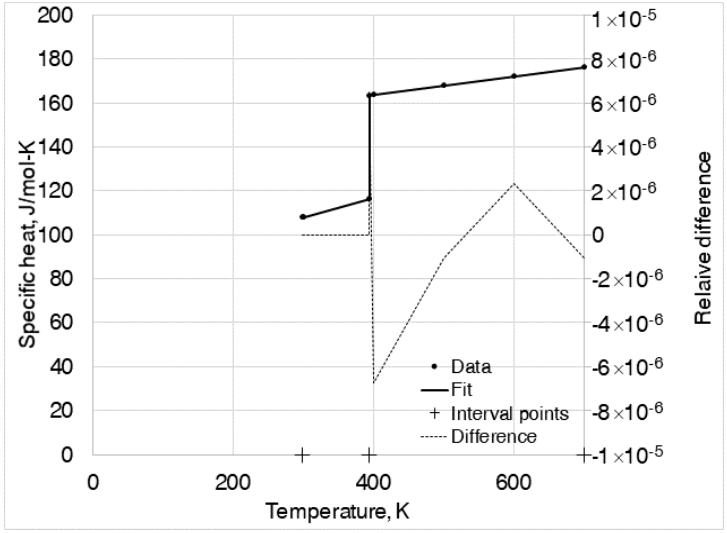


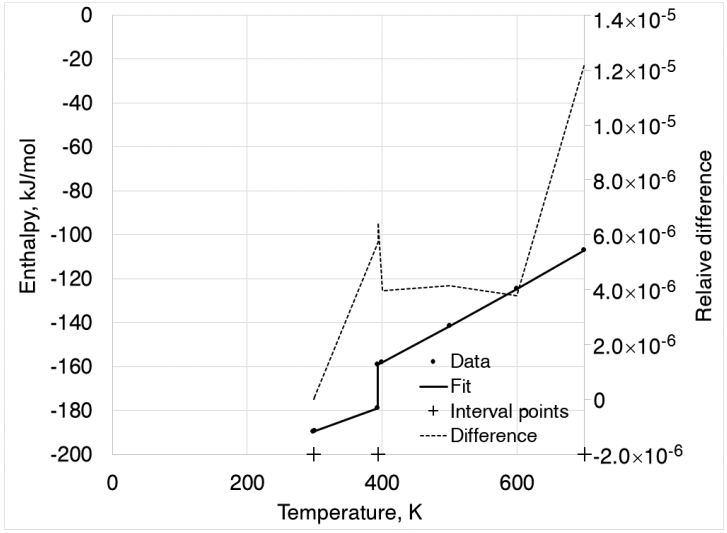


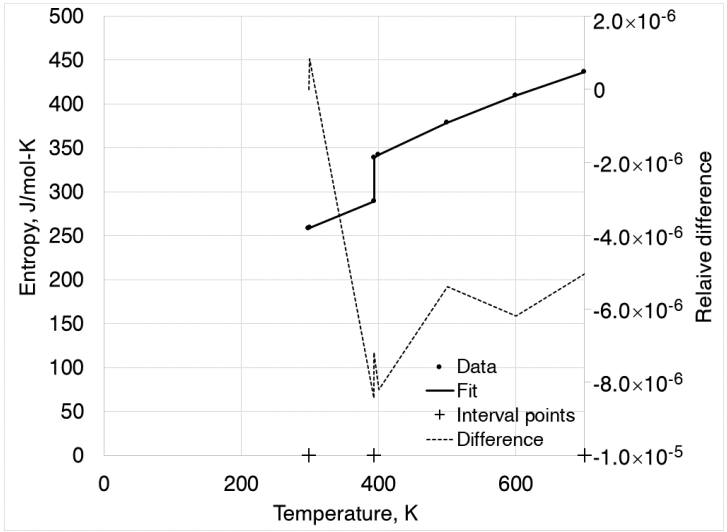


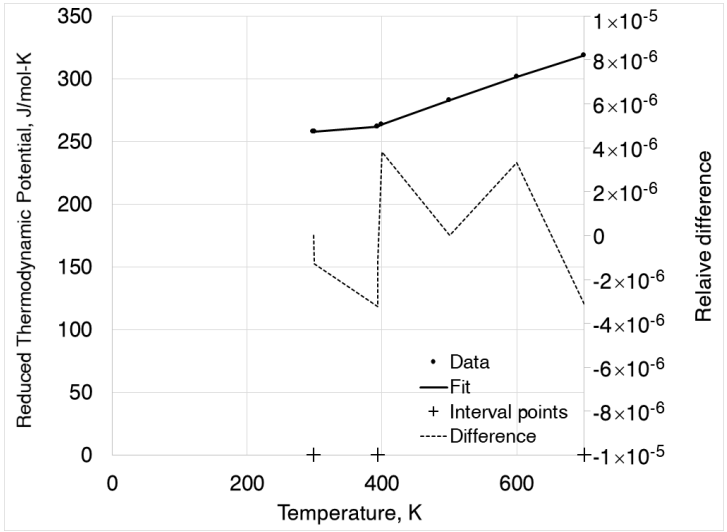
Tetraiodosilane (SiI_4) – I4Si1(cr,l)

MW = 535.7035 g/mol – $\Delta_f H^o$(298.15 K) = -189.535 kJ/mol – S^o(298.15 K) = 258.153 J/mol-K – p^o = 0.1 MPa



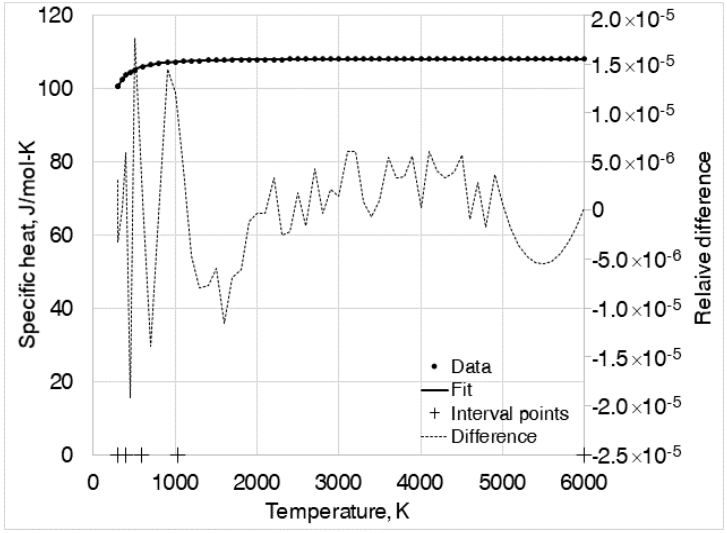


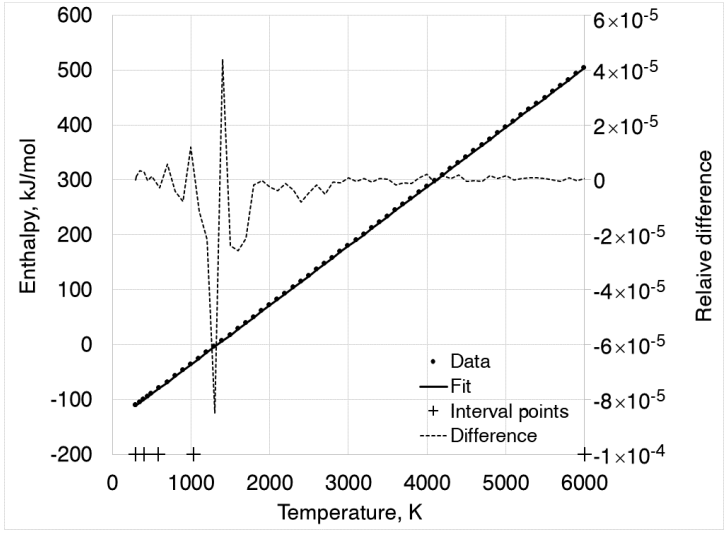


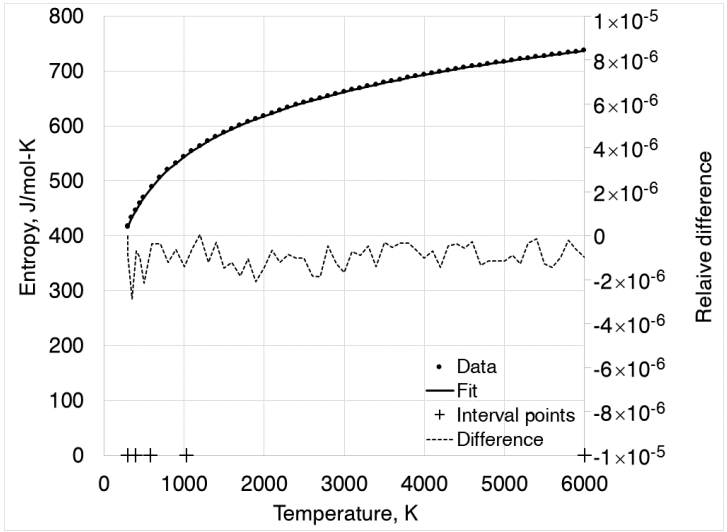


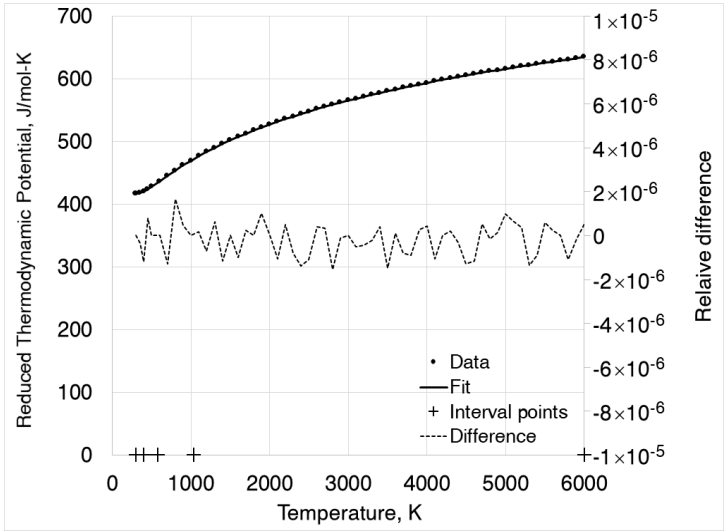
Tetraiodosilane (SiI_4) – I4Si1(g)

MW = 535.7035 g/mol – $\Delta_f H^o$(298.15 K) = -110.458 kJ/mol – S^o(298.15 K) = 416.449 J/mol-K – p^o = 0.1 MPa



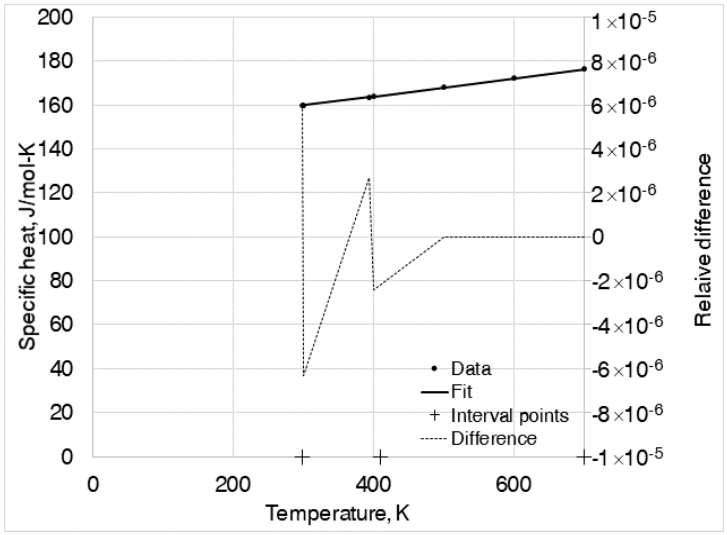


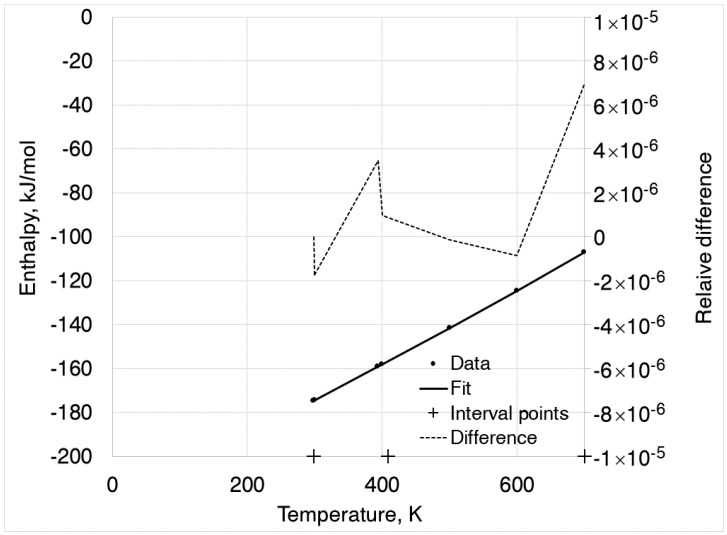


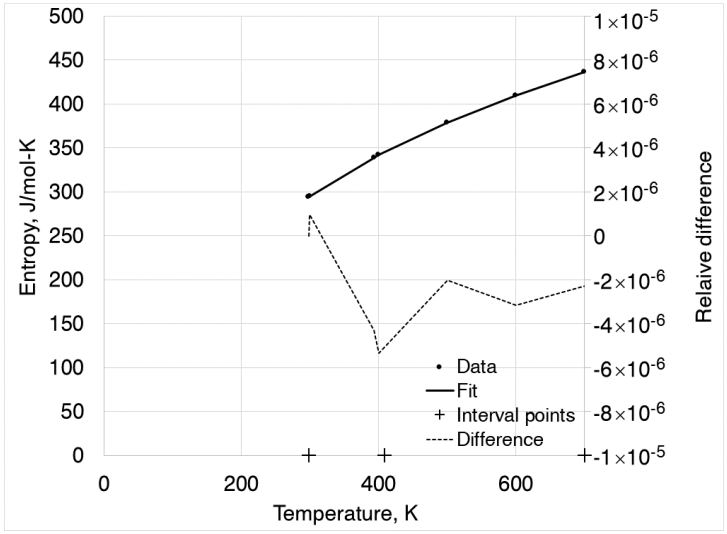


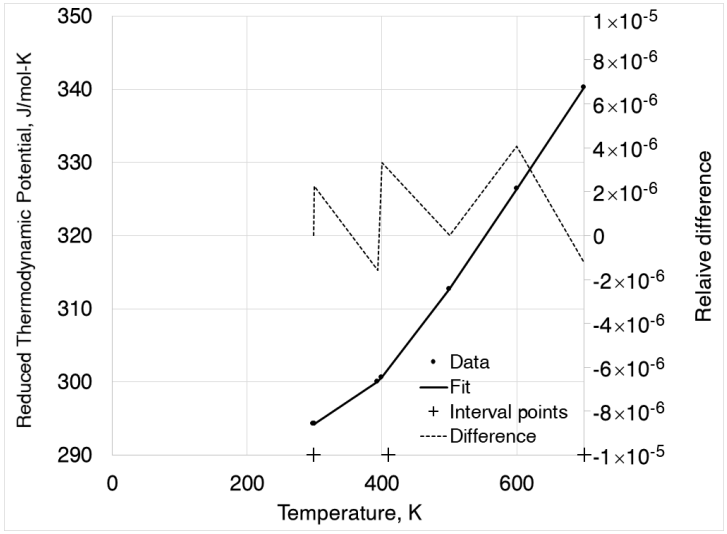
Tetraiodosilane (SiI_4) – I4Si1(l)

MW = 535.7035 g/mol – $\Delta_f H^o$(298.15 K) = -174.603 kJ/mol – S^o(298.15 K) = 294.309 J/mol-K – p^o = 0.1 MPa



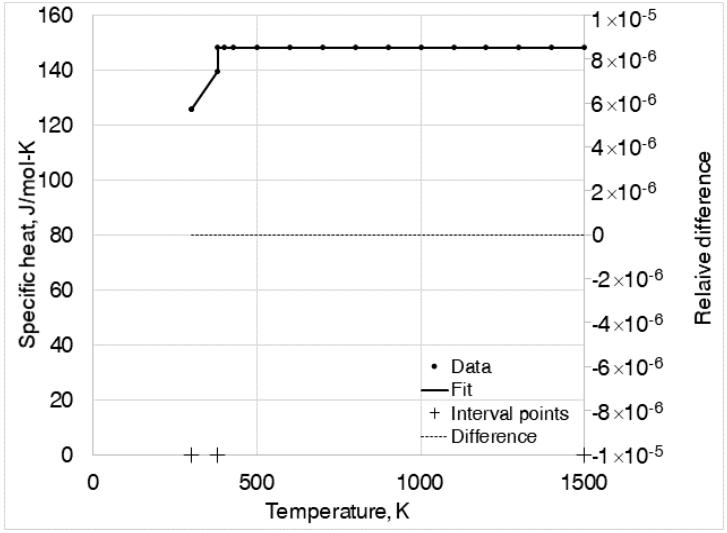


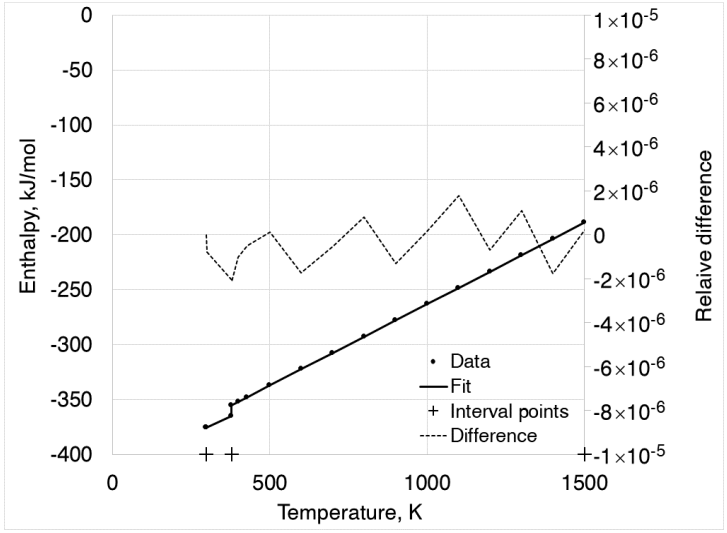


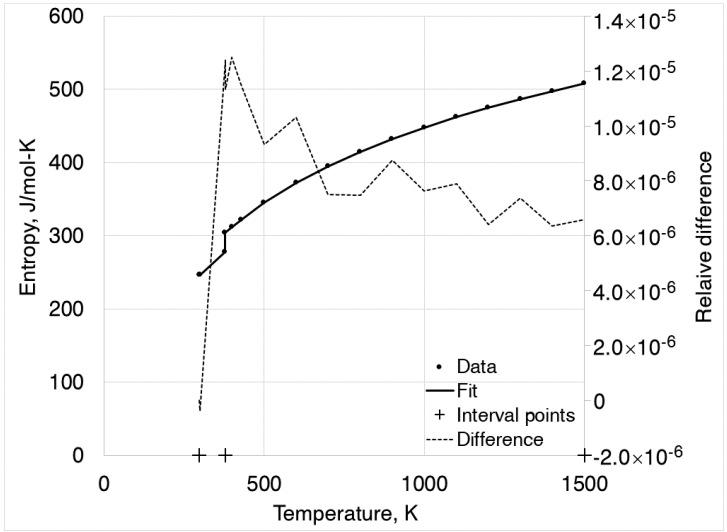


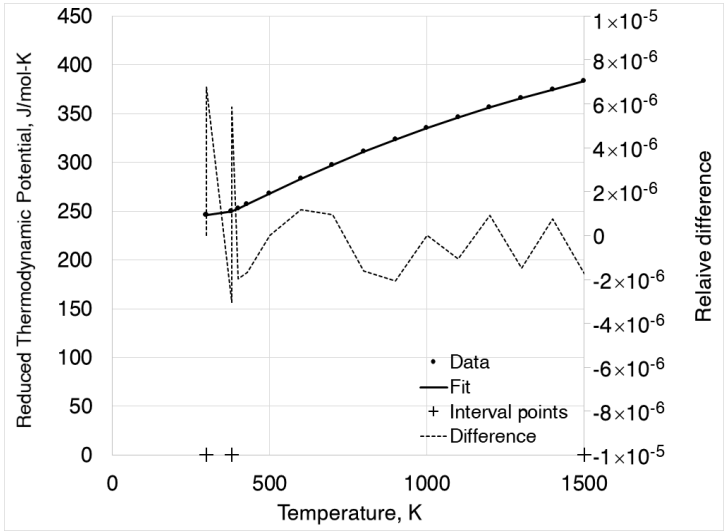
Titanium Iodide (TiI_4) – I4Ti1(cr)

MW = 555.5 g/mol – $\Delta_f H^o$(298.15 K) = -375.723 kJ/mol – S^o(298.15 K) = 246.15 J/mol-K – p^o = 0.1 MPa



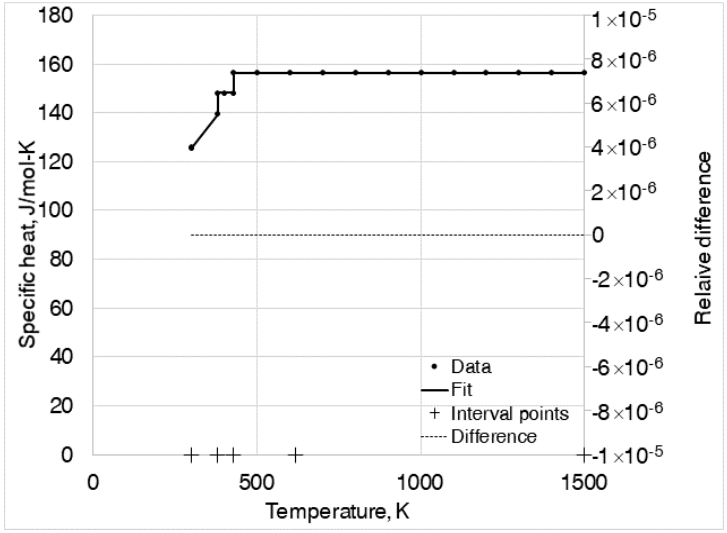


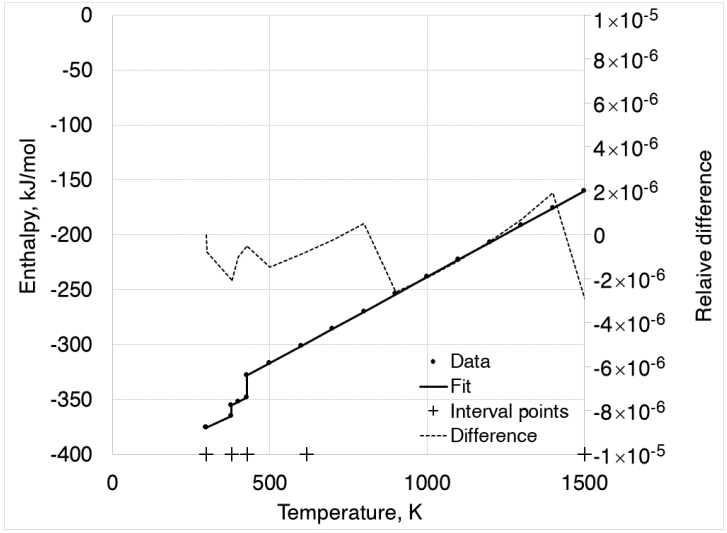


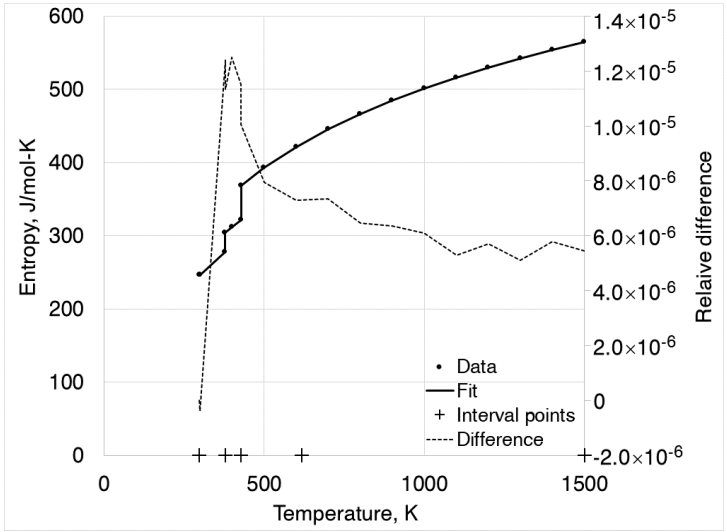


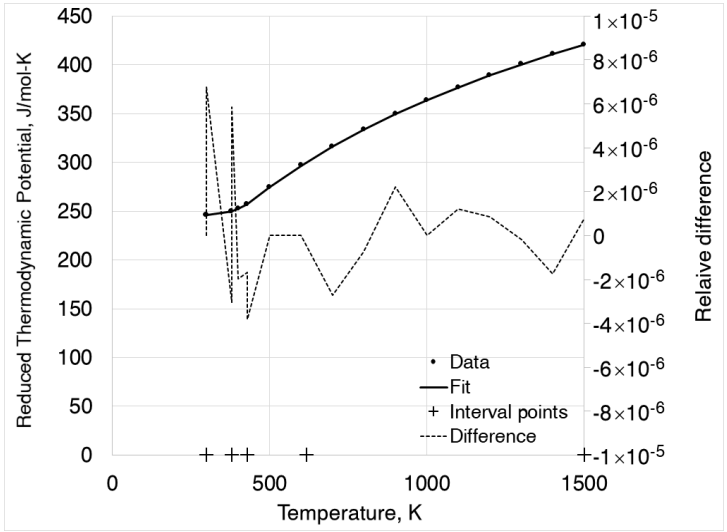
Titanium Iodide (TiI_4) – I4Ti1(cr,l)

MW = 555.5 g/mol – $\Delta_f H^o$(298.15 K) = -375.723 kJ/mol – S^o(298.15 K) = 246.15 J/mol-K – p^o = 0.1 MPa



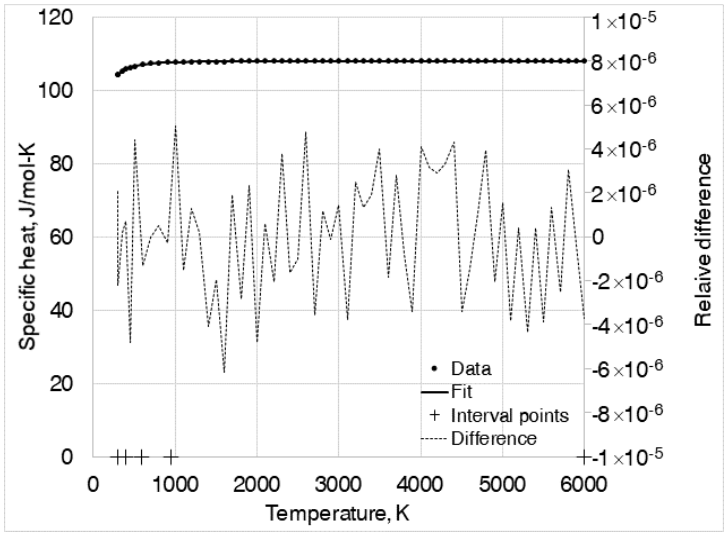


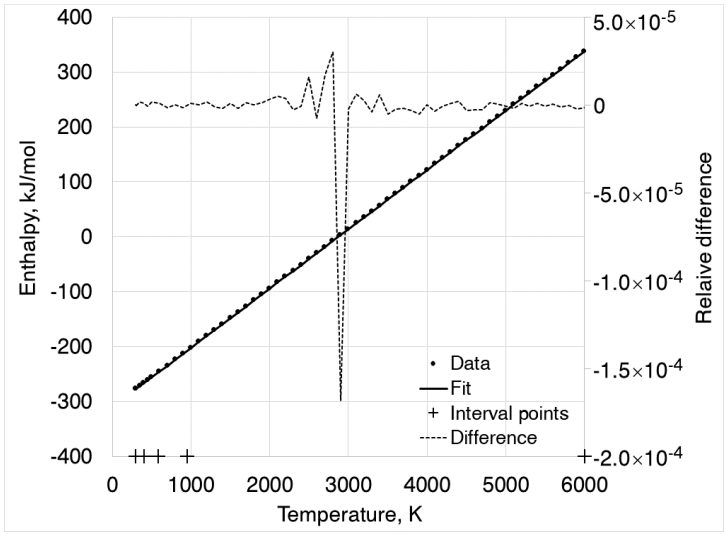


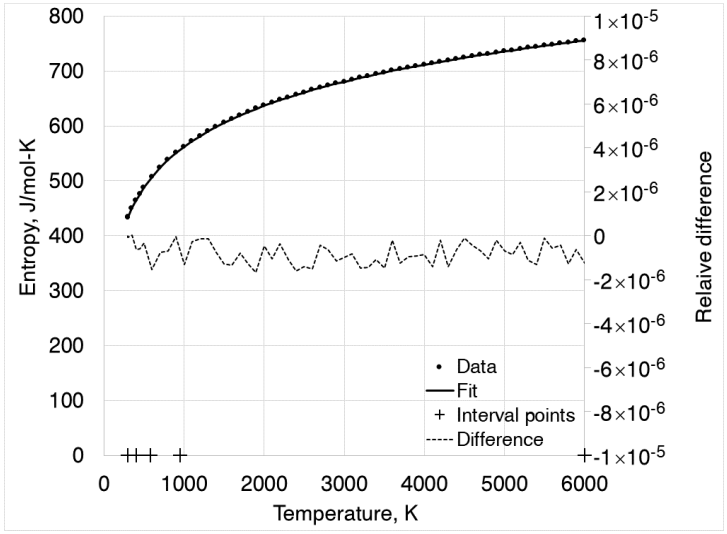


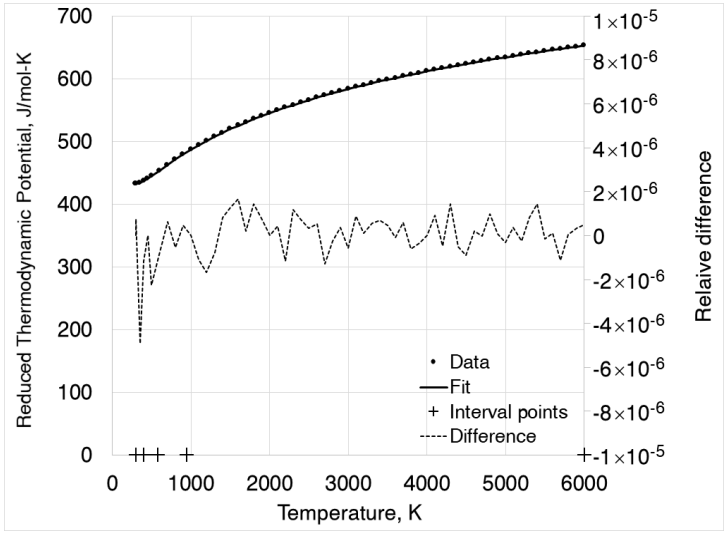
Titanium Iodide (TiI_4) – I4Ti1(g)

MW = 555.5 g/mol – $\Delta_f H^o$(298.15 K) = -277.274 kJ/mol – S^o(298.15 K) = 433.072 J/mol-K – p^o = 0.1 MPa



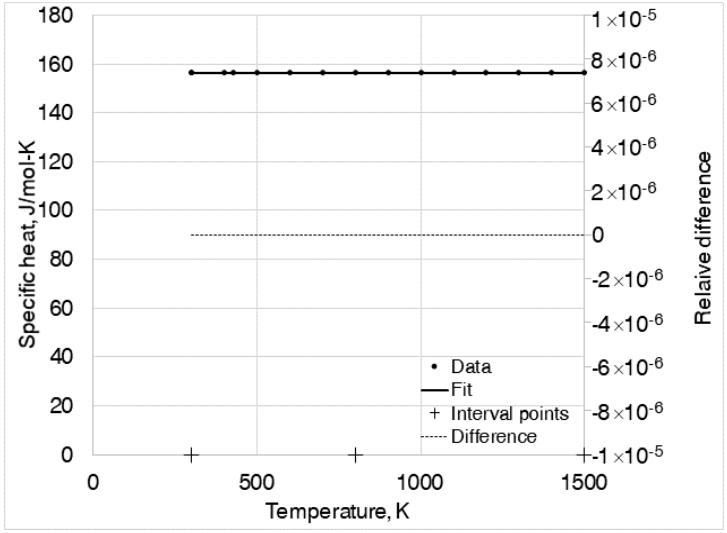


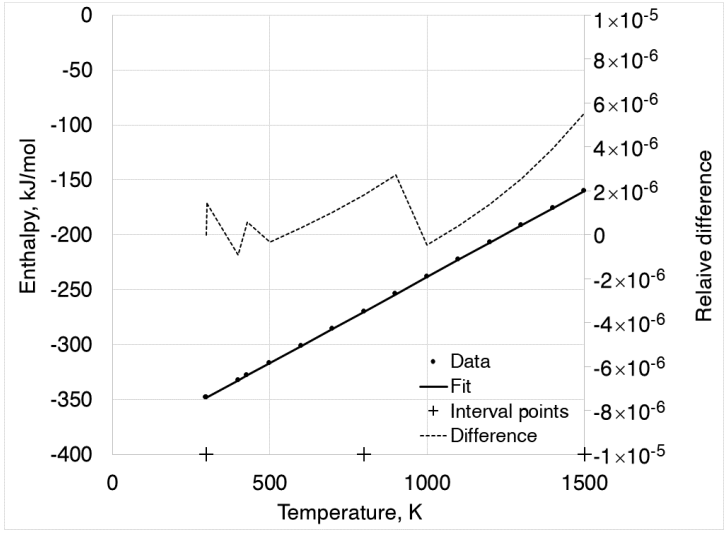


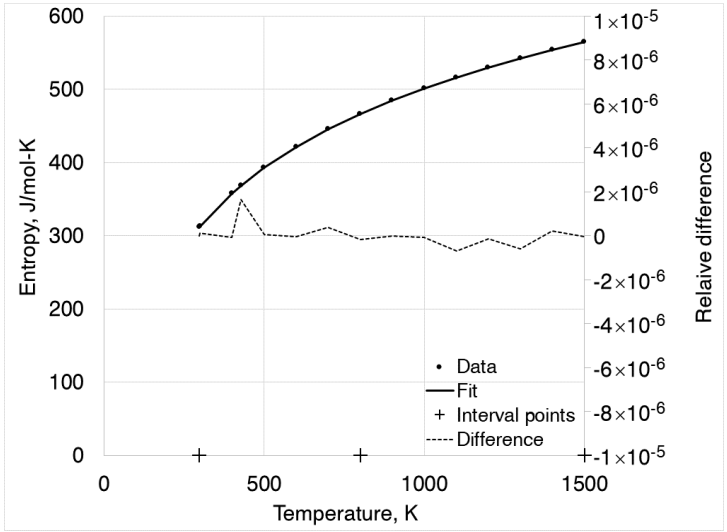


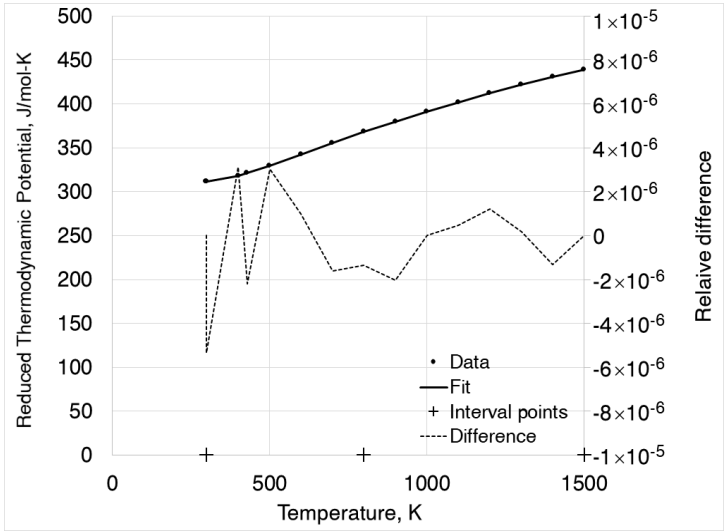
Titanium Iodide (TiI_4) – I4Ti1(l)

MW = 555.5 g/mol – $\Delta_f H^o$(298.15 K) = -348.314 kJ/mol – S^o(298.15 K) = 311.841 J/mol-K – p^o = 0.1 MPa



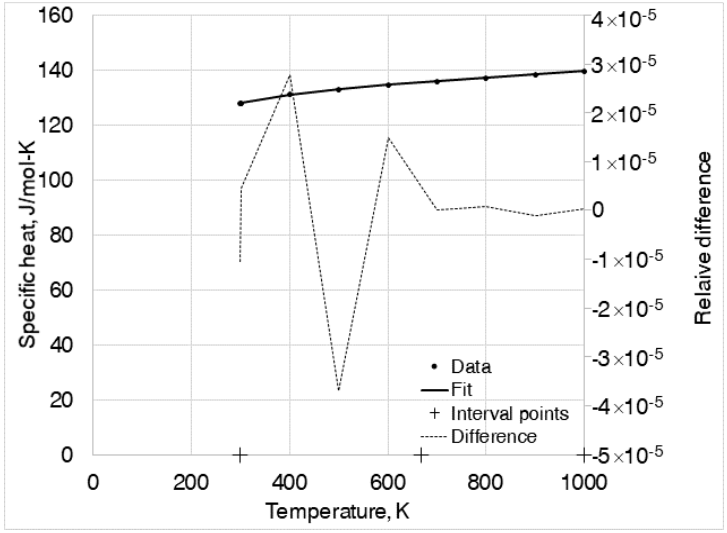


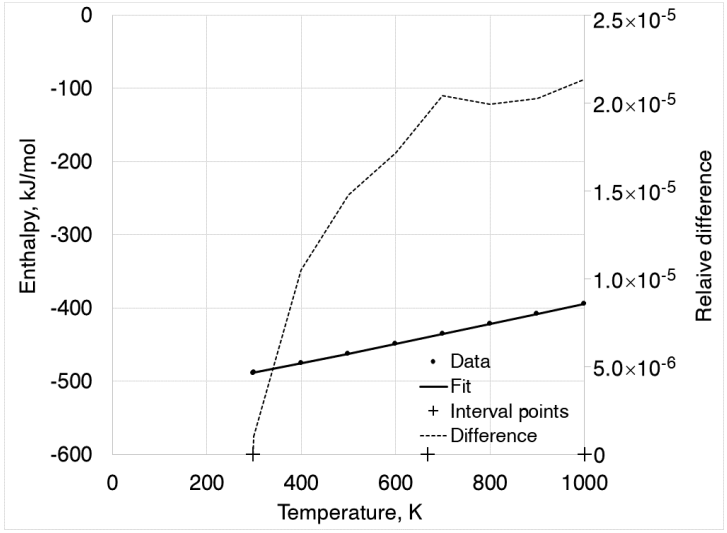


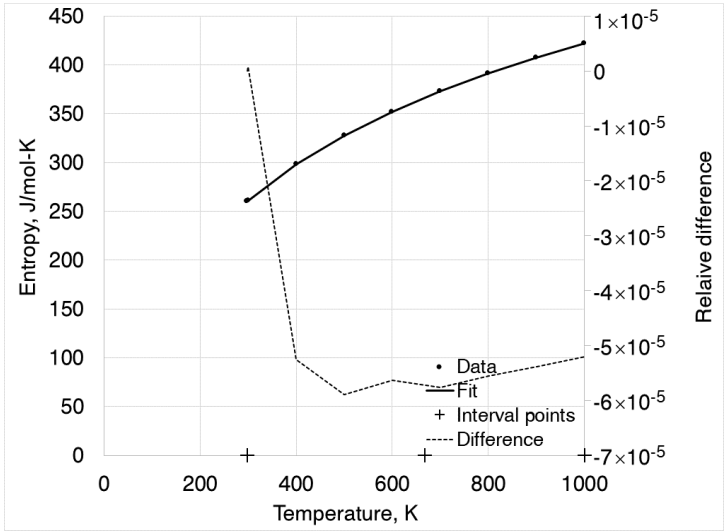


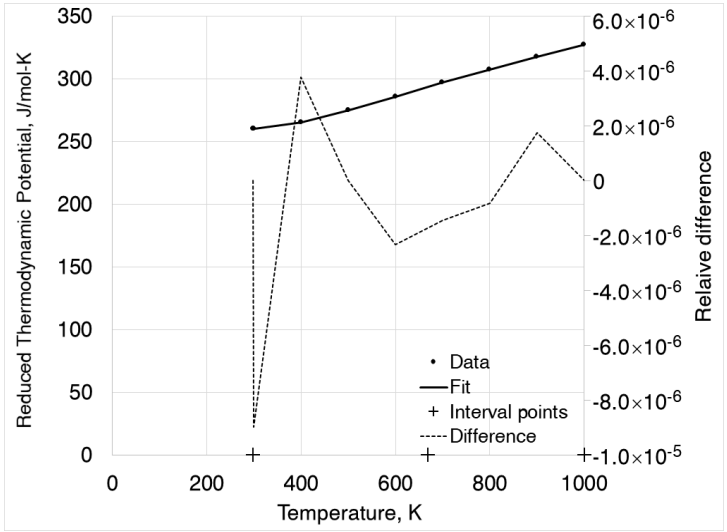
Zirconium Iodide (ZrI_4) – I4Zr1(cr)

MW = 598.84 g/mol – $\Delta_f H^o$(298.15 K) = -488.691 kJ/mol – S^o(298.15 K) = 260.27 J/mol-K – p^o = 0.1 MPa



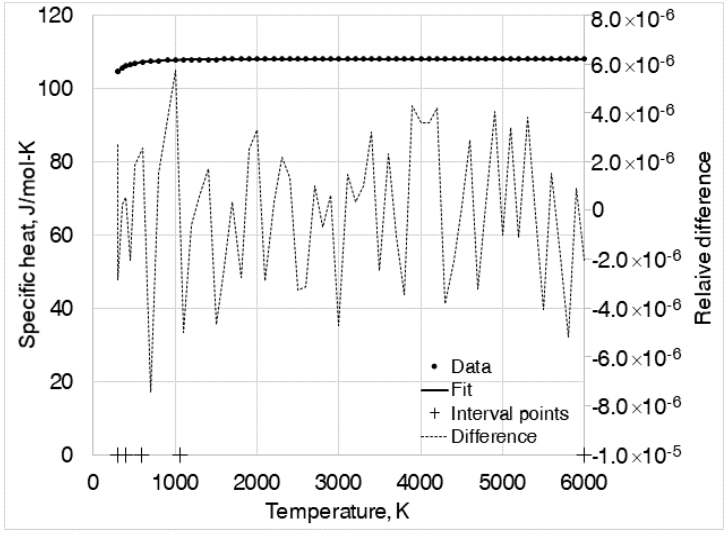


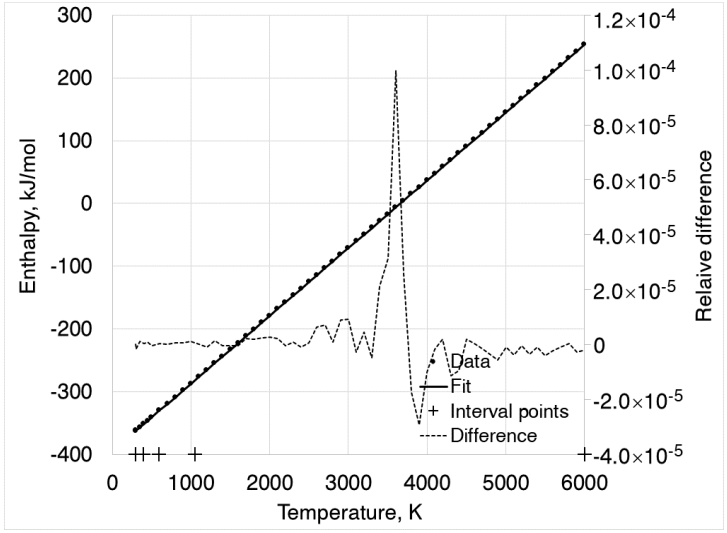


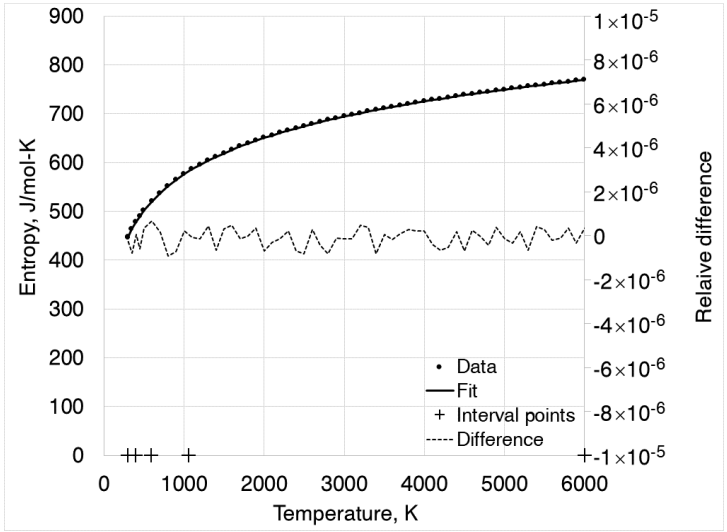


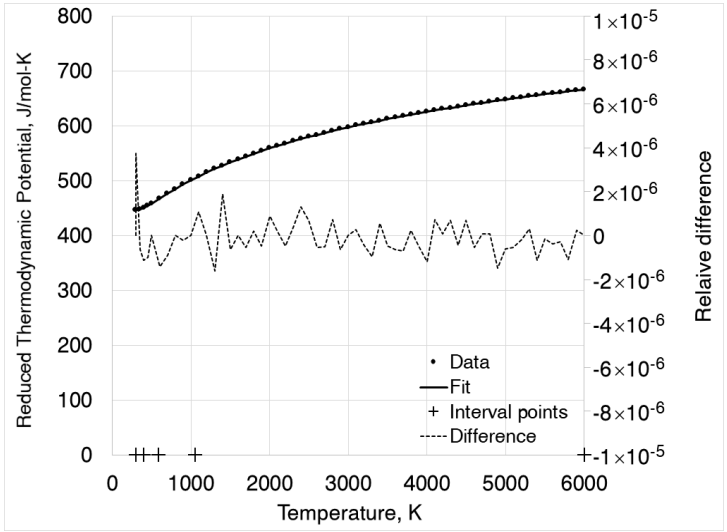
Zirconium Iodide (ZrI_4) – I4Zr1(g)

MW = 598.84 g/mol – $\Delta_f H^o$(298.15 K) = -362.334 kJ/mol – S^o(298.15 K) = 446.563 J/mol-K – p^o = 0.1 MPa



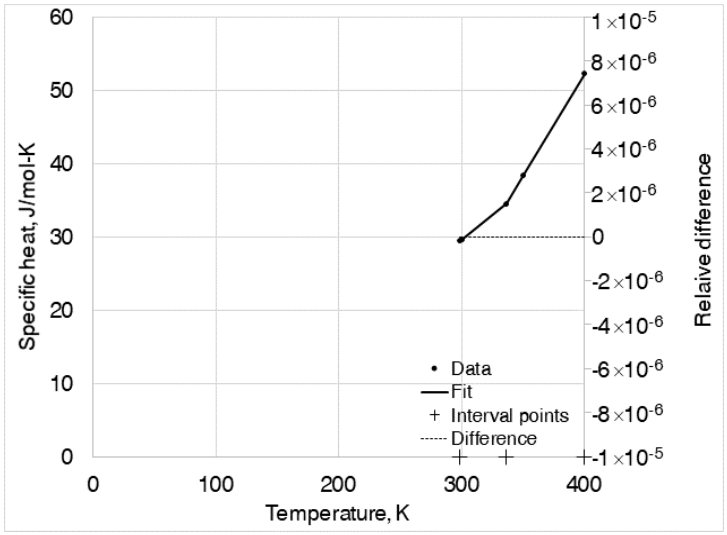


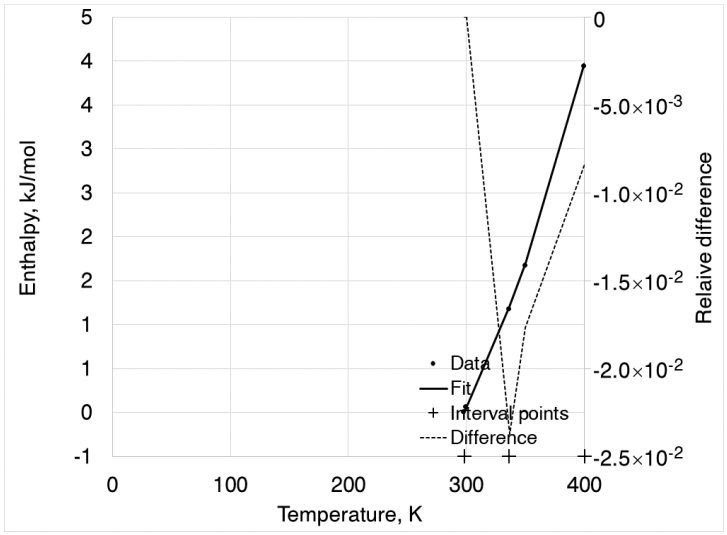


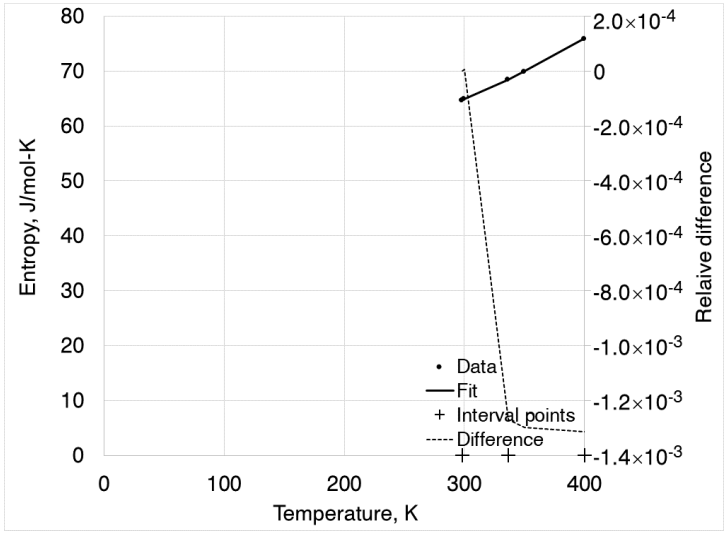


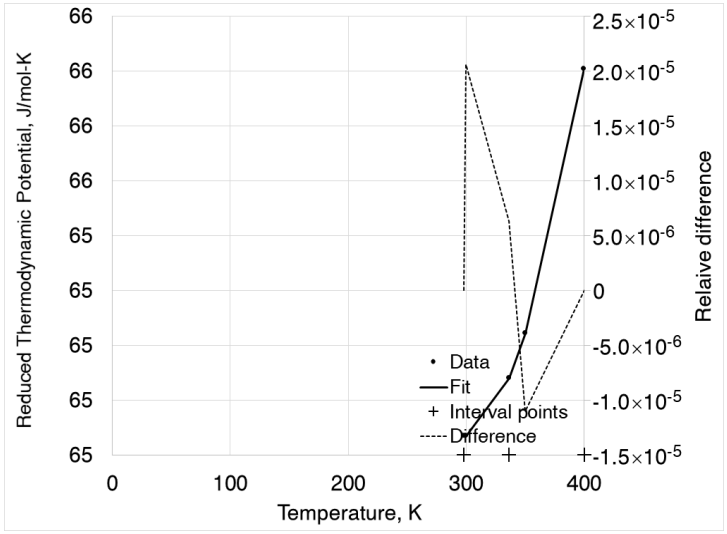
Potassium (K) – K1(cr)

MW = 39.0983 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 64.67 J/mol-K – p^o = 0.1 MPa



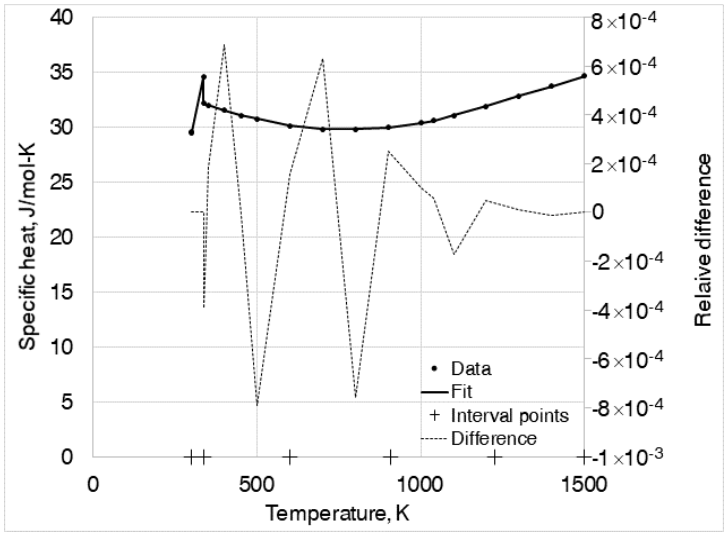


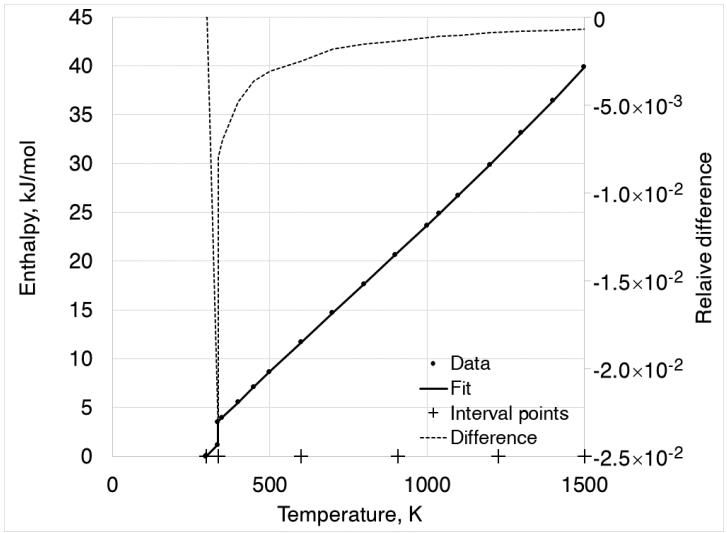


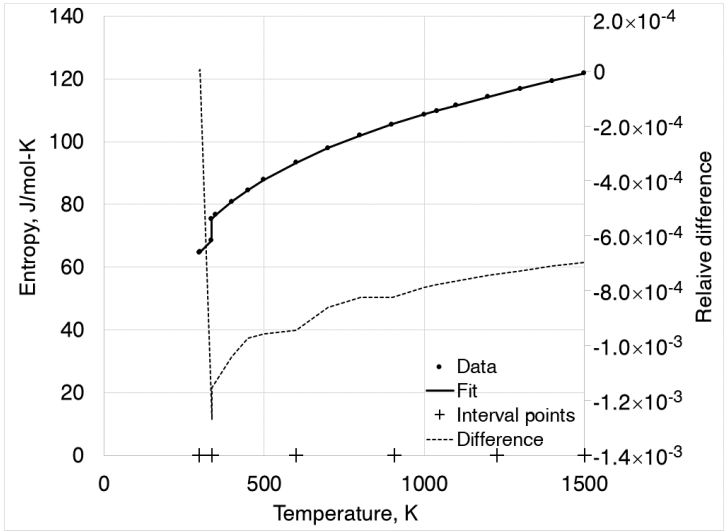


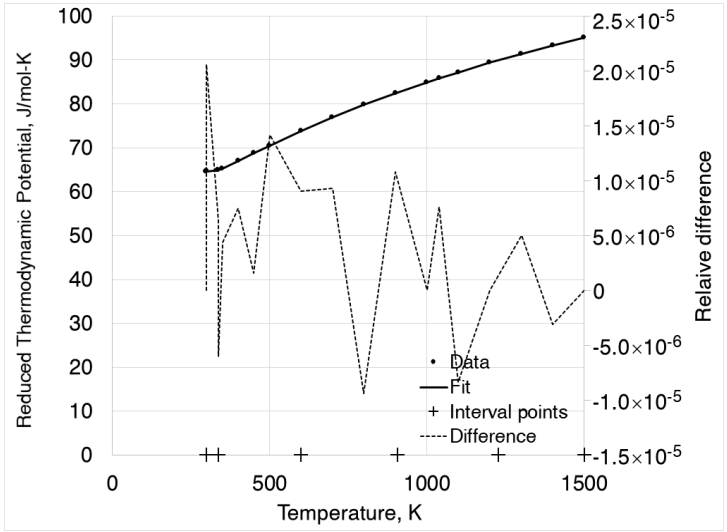
Potassium (K) – K1(cr,l)

MW = 39.0983 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 64.67 J/mol-K – p^o = 0.1 MPa



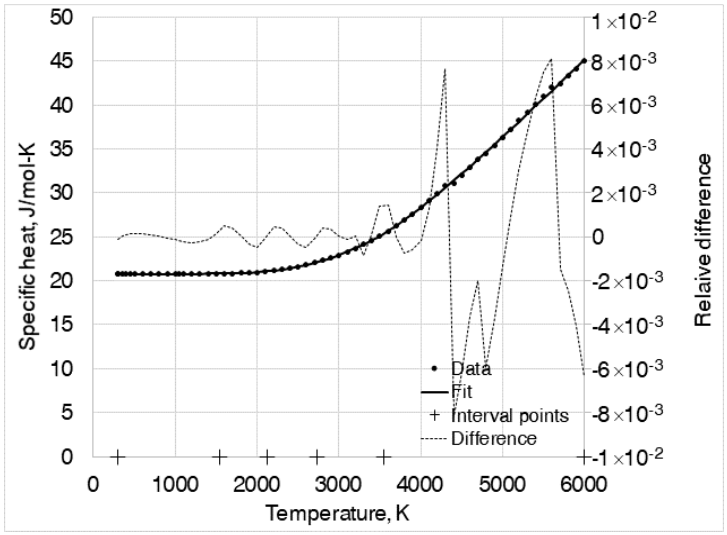


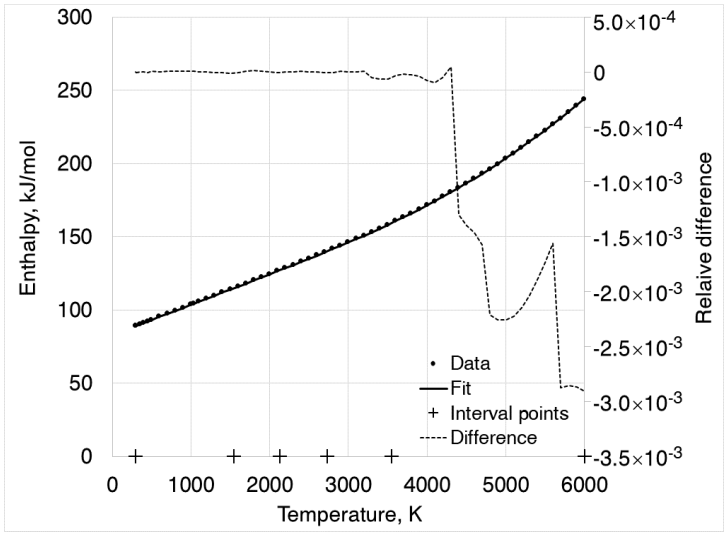


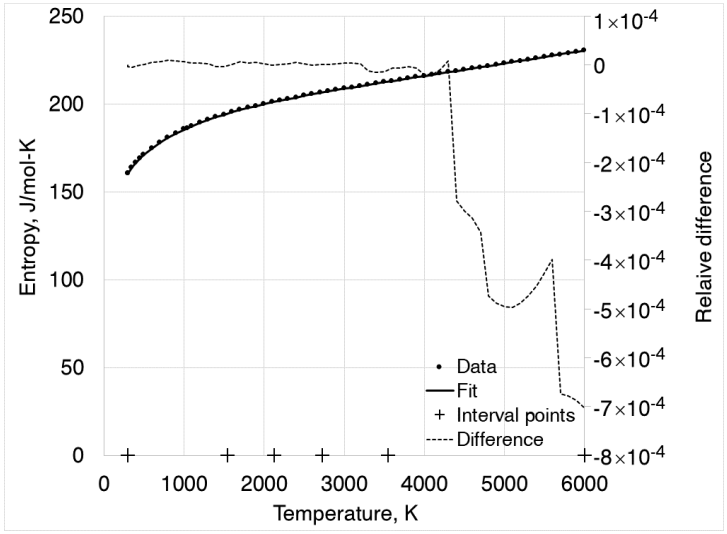


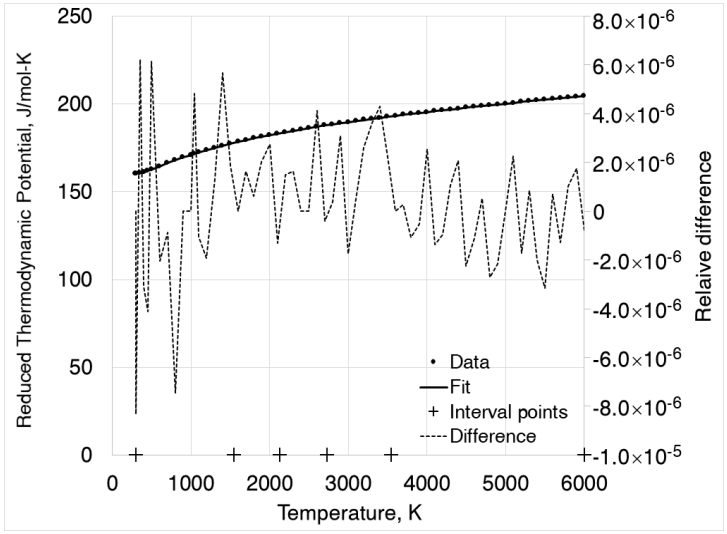
Potassium (K) – K1(g)

MW = 39.0983 g/mol – $\Delta_f H^o$(298.15 K) = 89 kJ/mol – S^o(298.15 K) = 160.34 J/mol-K – p^o = 0.1 MPa



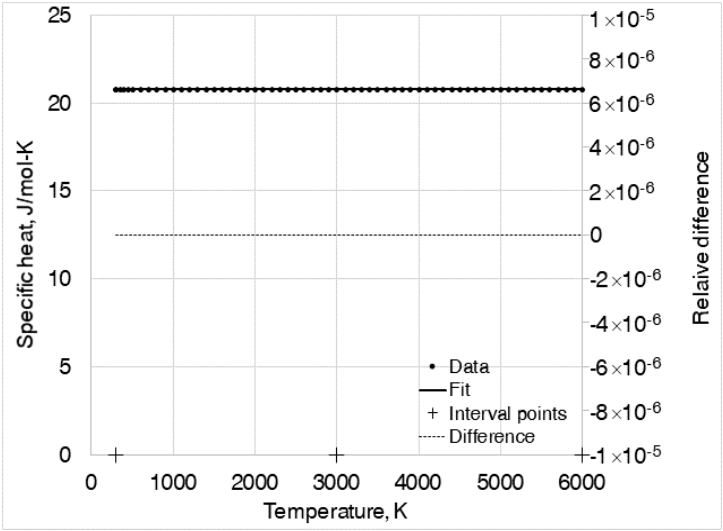


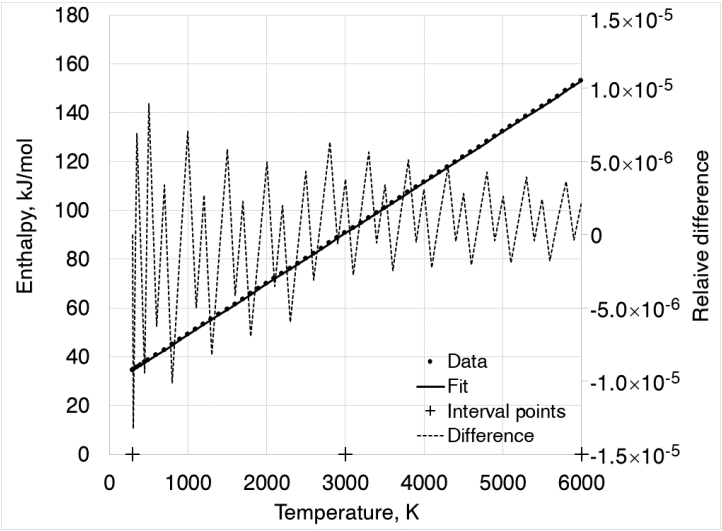


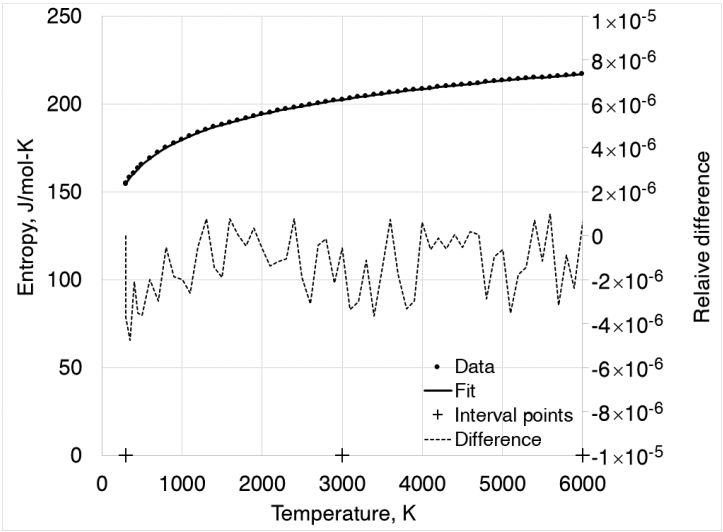


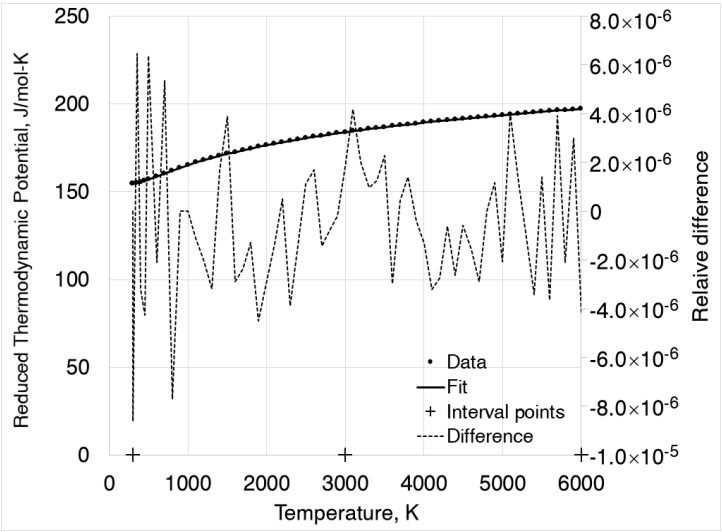
Potassium, Ion (K^-) – K1-(g)

MW = 39.0988 g/mol – $\Delta_f H^o$(298.15 K) = 34.419 kJ/mol – S^o(298.15 K) = 154.577 J/mol-K – p^o = 0.1 MPa



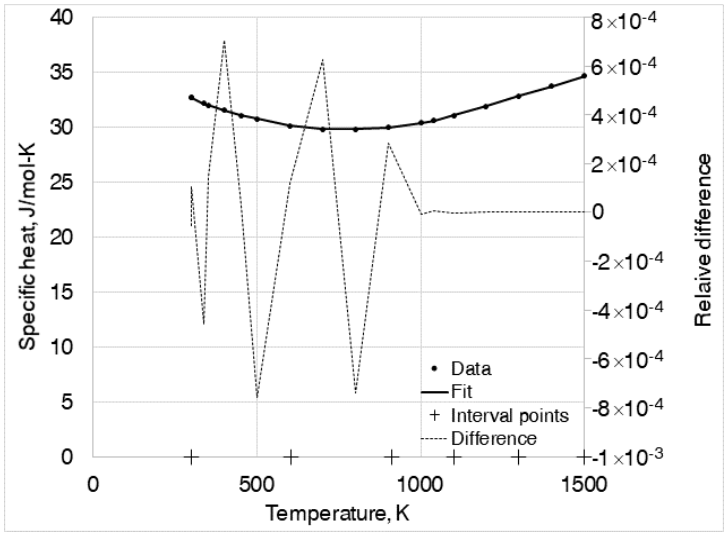


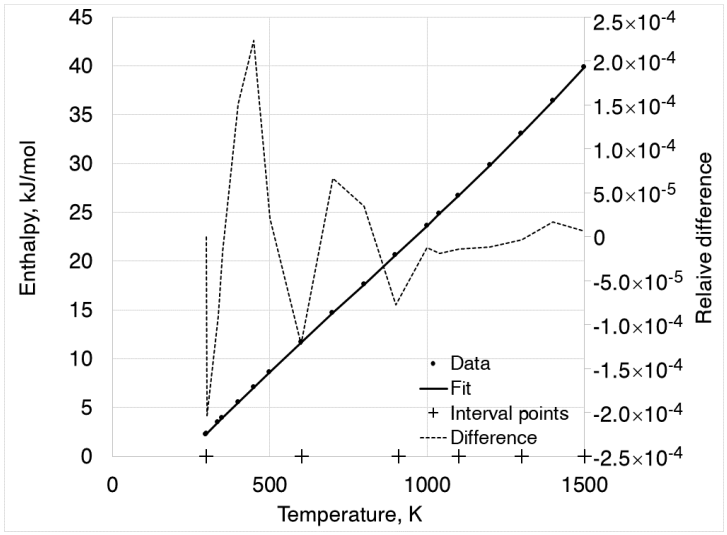


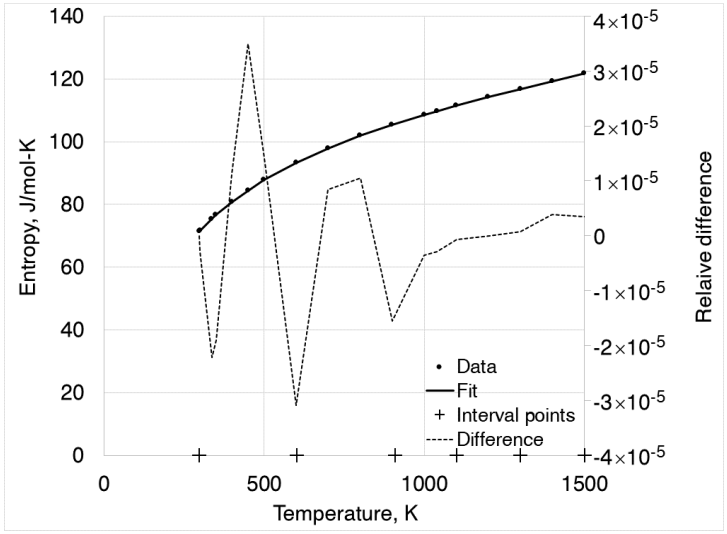


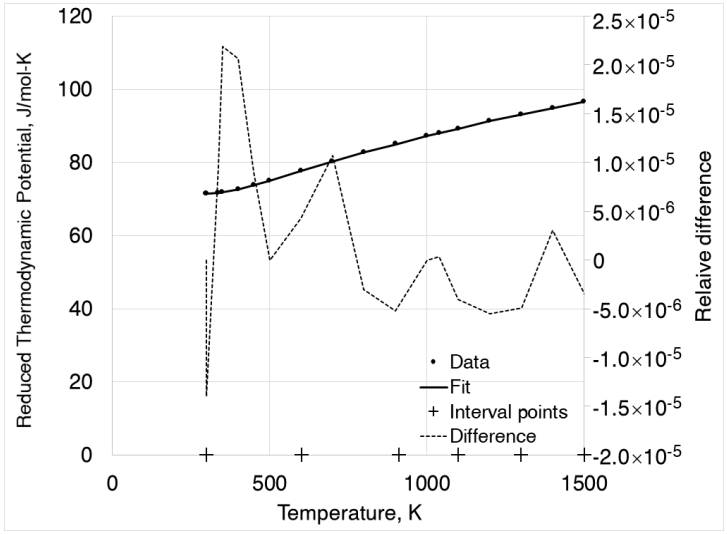
Potassium (K) – K1(l)

MW = 39.0983 g/mol – $\Delta_f H^o$(298.15 K) = 2.27 kJ/mol – S^o(298.15 K) = 71.403 J/mol-K – p^o = 0.1 MPa



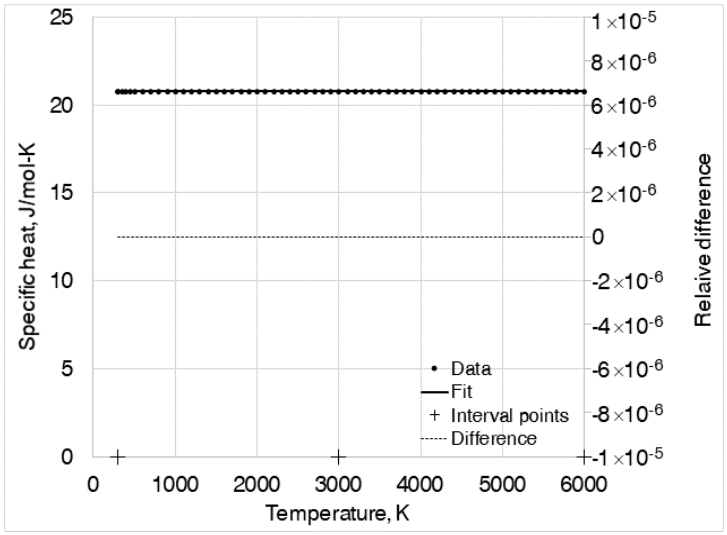


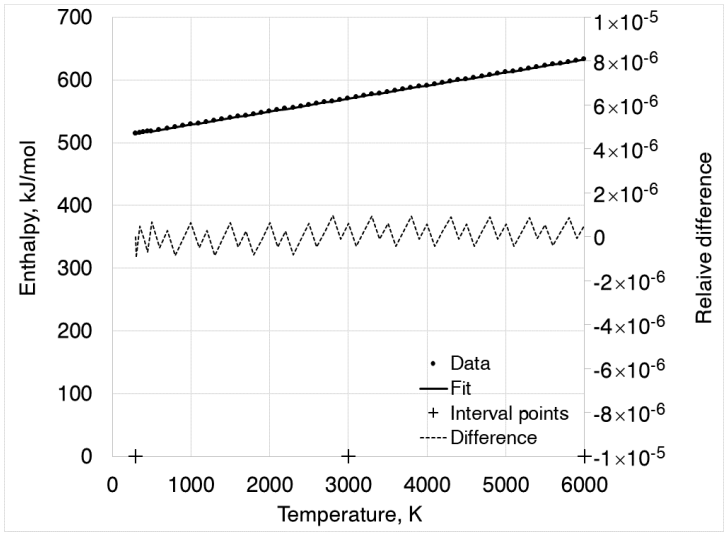


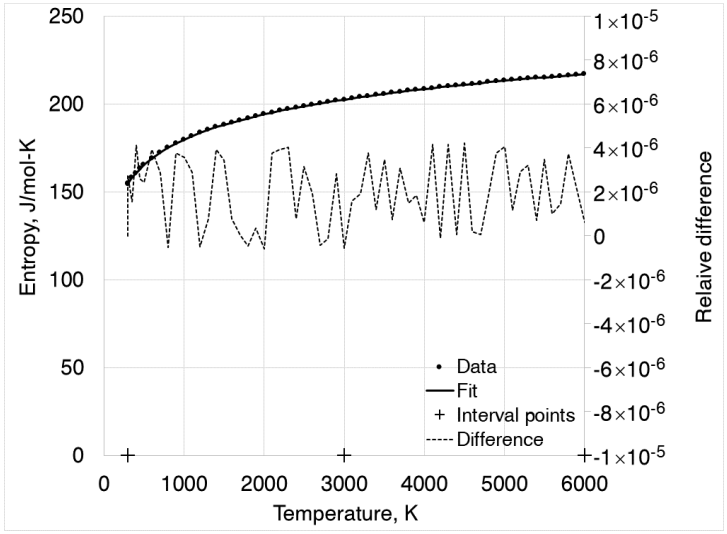


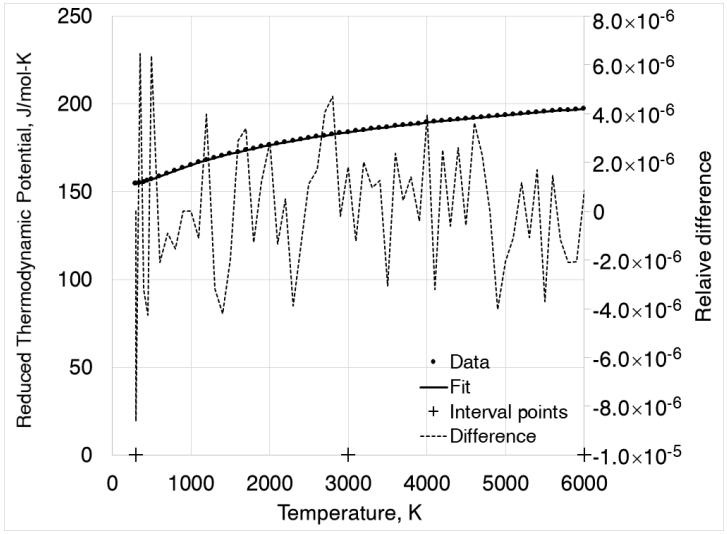
Potassium, Ion (K^+) – K1+(g)

MW = 39.0978 g/mol – $\Delta_f H^o$(298.15 K) = 514.007 kJ/mol – S^o(298.15 K) = 154.576 J/mol-K – p^o = 0.1 MPa



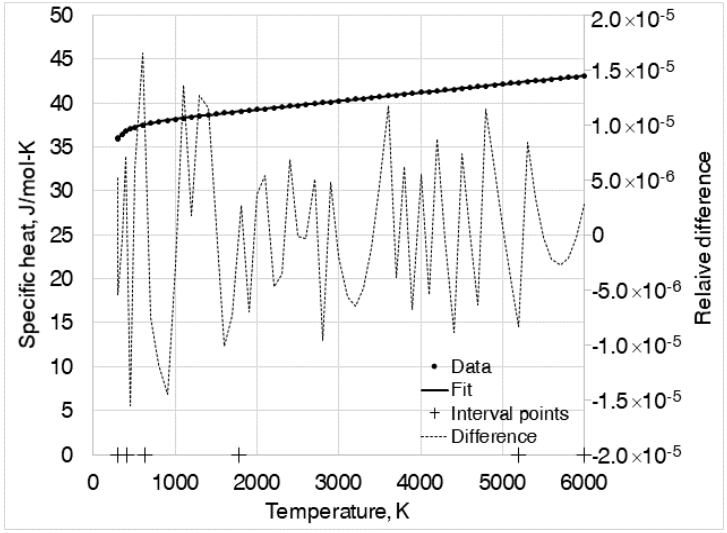


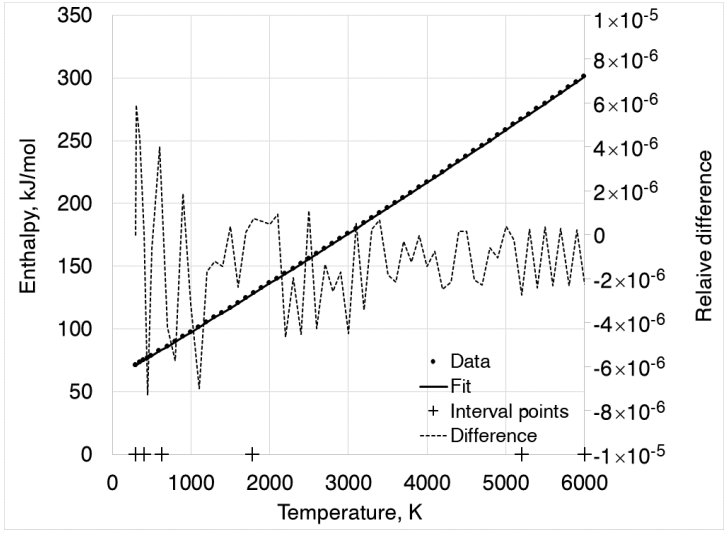


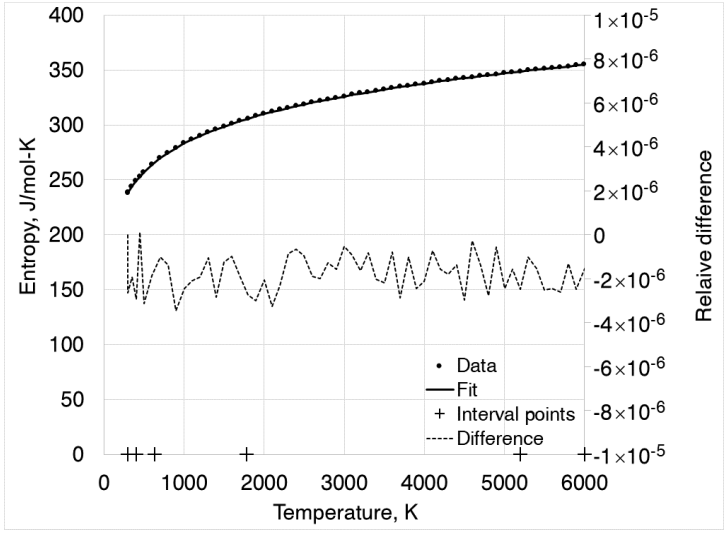


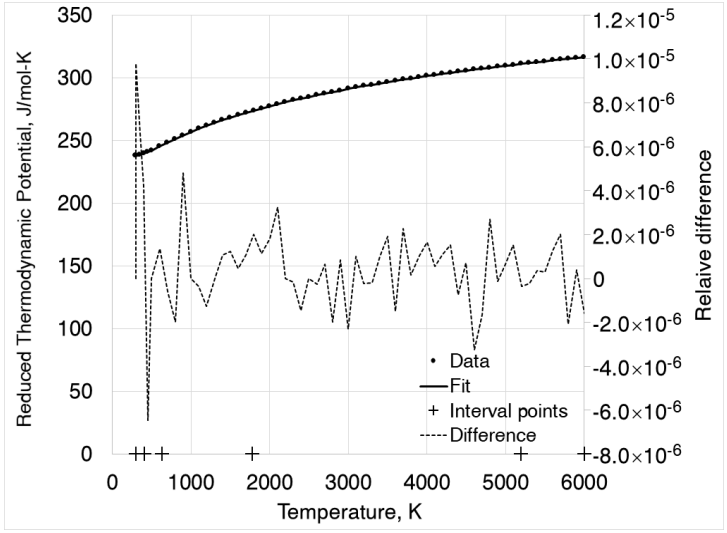
Potassium Oxide (KO) – K1O1(g)

MW = 55.0977 g/mol – $\Delta_f H^o$(298.15 K) = 71.128 kJ/mol – S^o(298.15 K) = 238.027 J/mol-K – p^o = 0.1 MPa



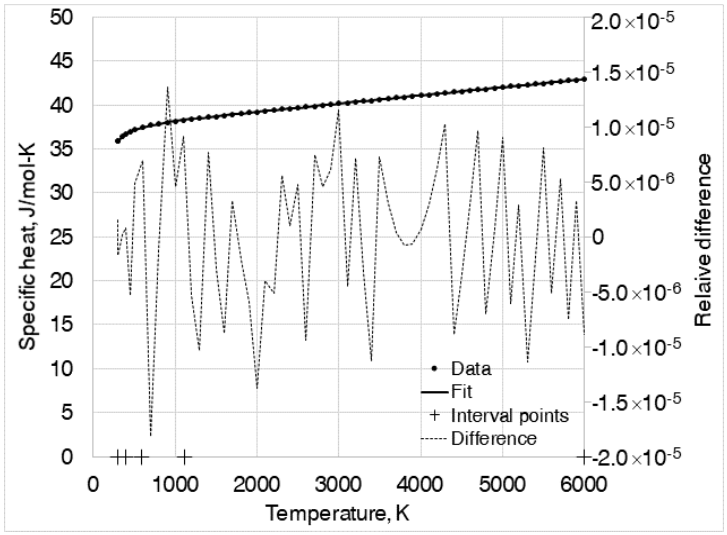


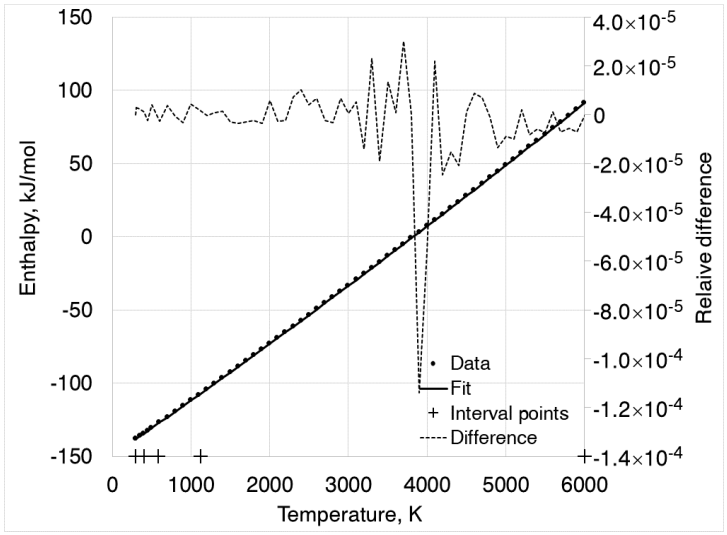


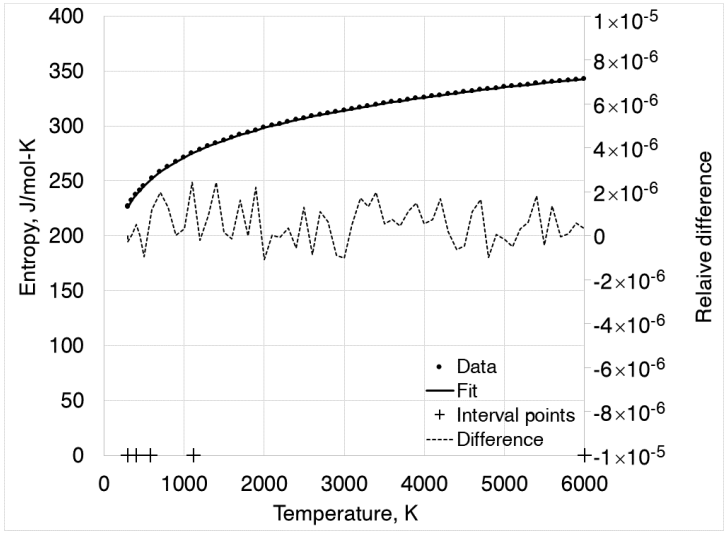


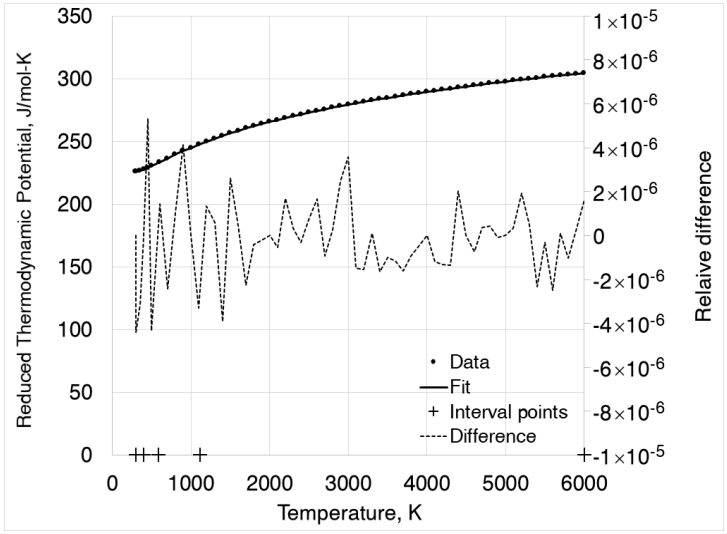
Potassium Oxide, Ion (KO^-) – K1O1-(g)

MW = 55.0982 g/mol – $\Delta_f H^o$(298.15 K) = -138.072 kJ/mol – S^o(298.15 K) = 226.308 J/mol-K – p^o = 0.1 MPa



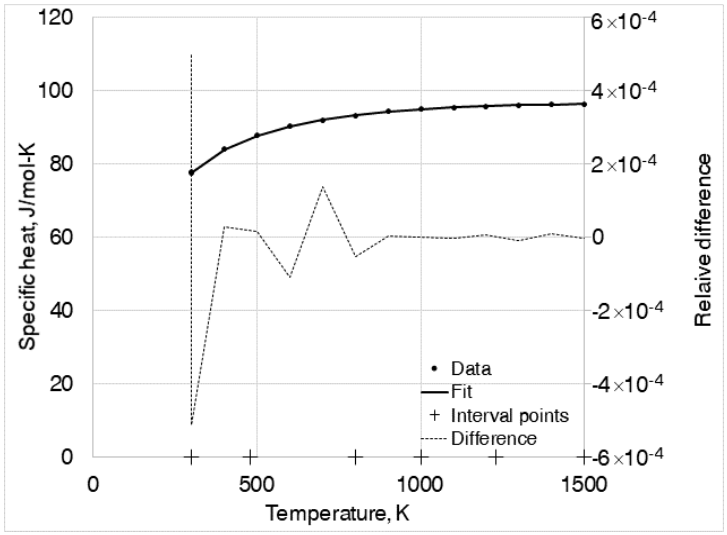


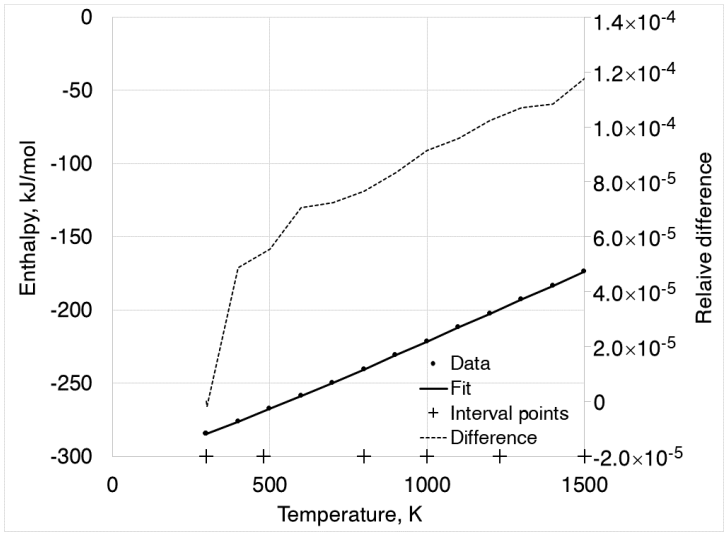


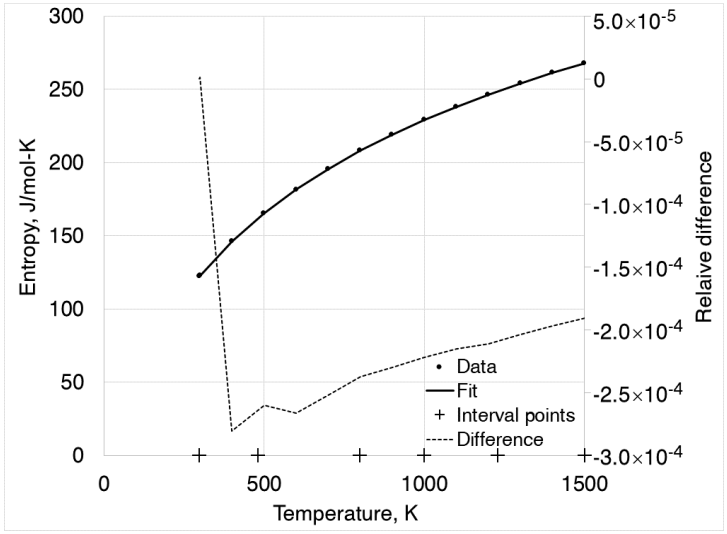


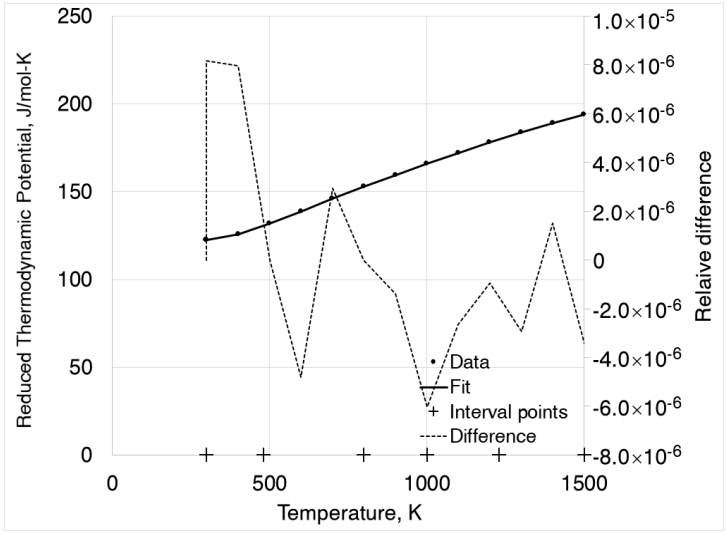
Potassium Superoxide (KO_2) – K1O2(cr)

MW = 71.0971 g/mol – $\Delta_f H^o$(298.15 K) = -284.512 kJ/mol – S^o(298.15 K) = 122.499 J/mol-K – p^o = 0.1 MPa



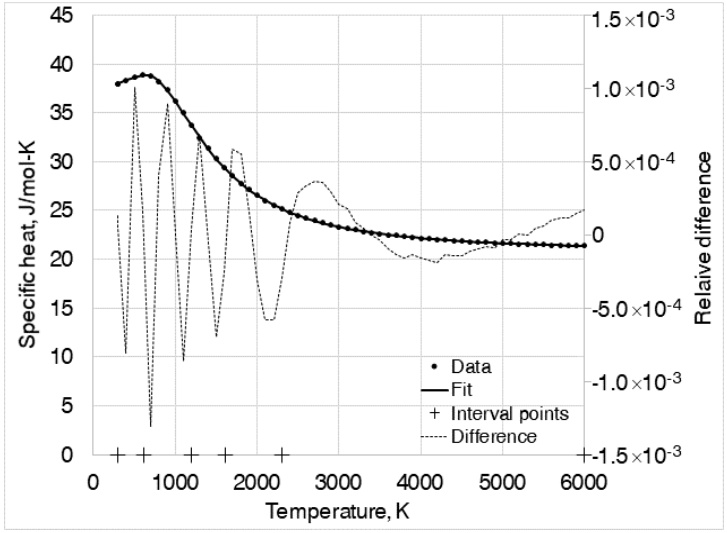


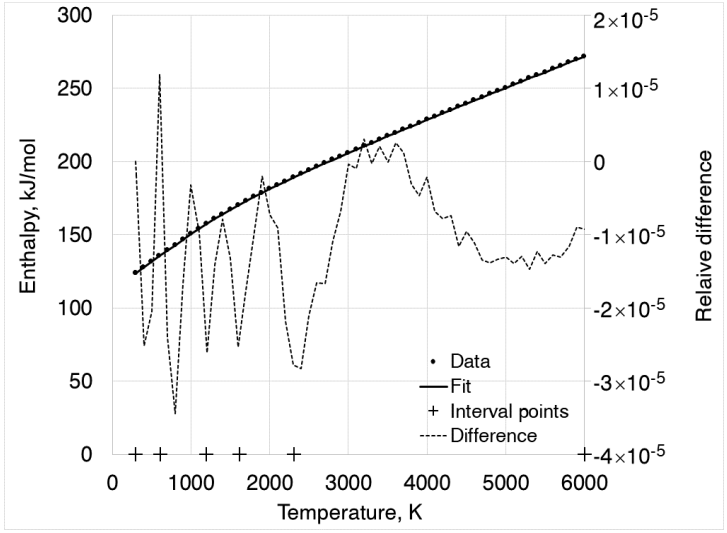


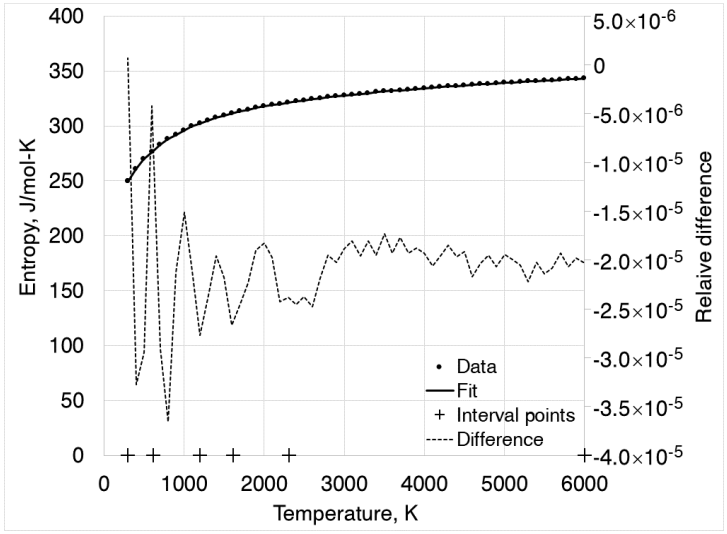


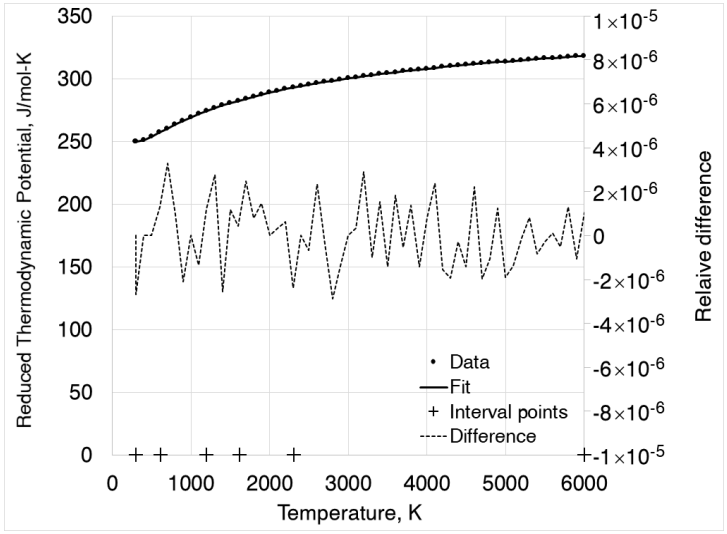
Potassium (K_2) – K2(g)

MW = 78.1966 g/mol – $\Delta_f H^o$(298.15 K) = 123.683 kJ/mol – S^o(298.15 K) = 249.69 J/mol-K – p^o = 0.1 MPa



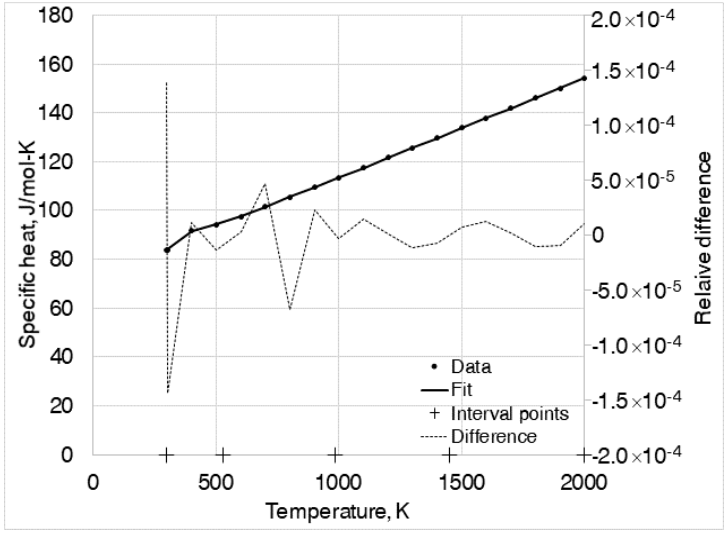


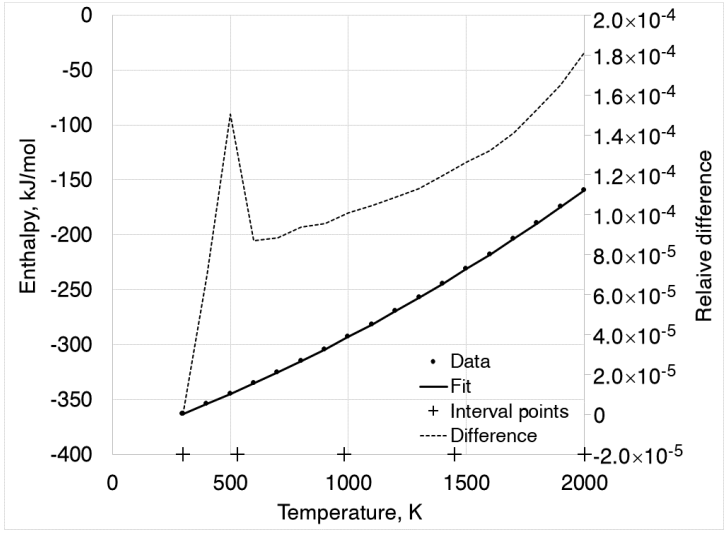


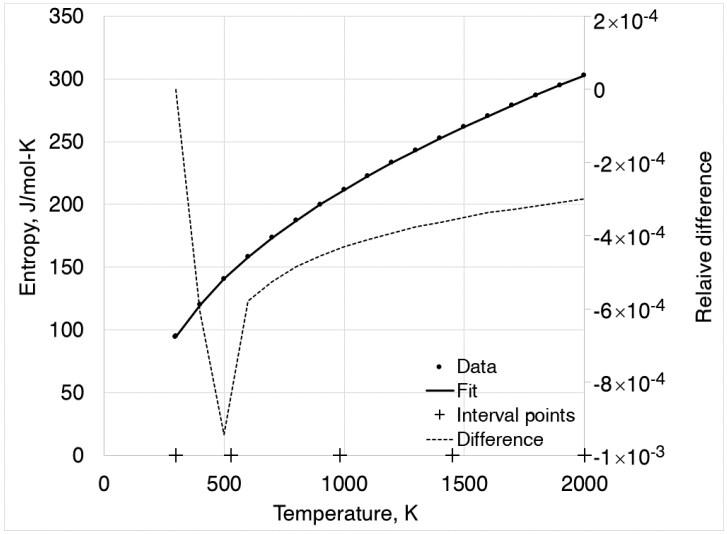


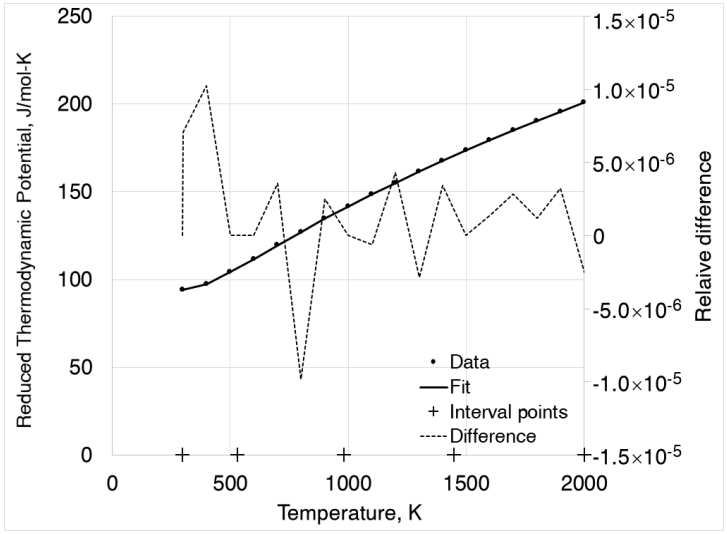
Potassium Oxide (K_2O) – K2O1(cr)

MW = 94.196 g/mol – $\Delta_f H^o$(298.15 K) = -363.171 kJ/mol – S^o(298.15 K) = 94.14 J/mol-K – p^o = 0.1 MPa



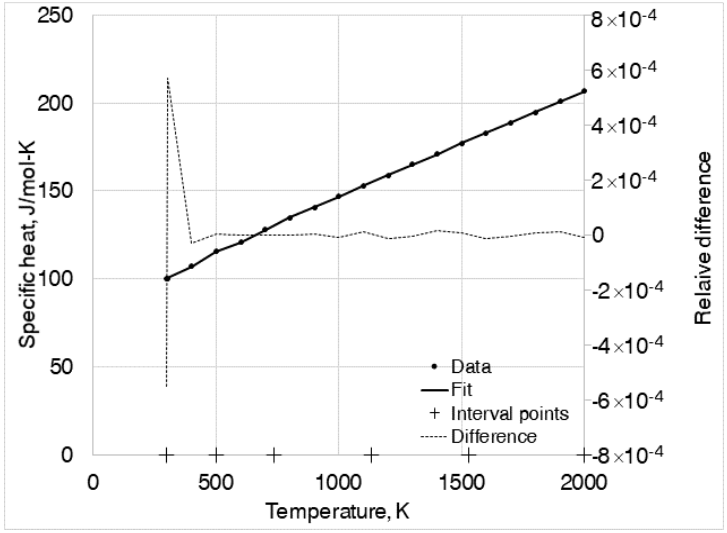


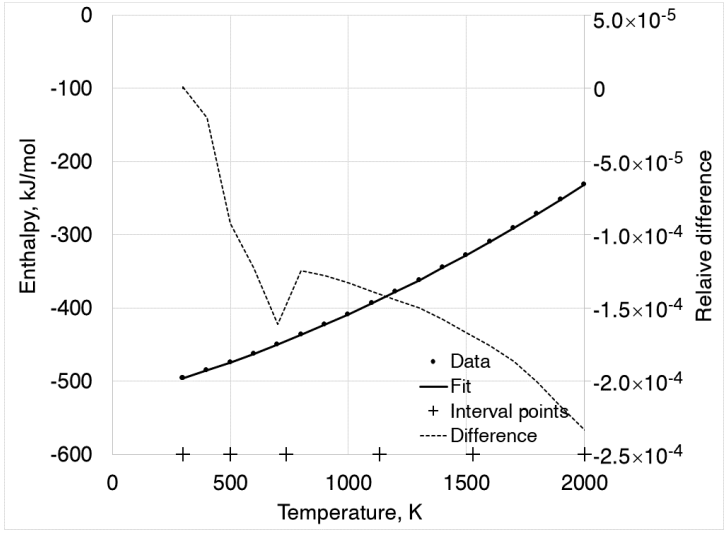


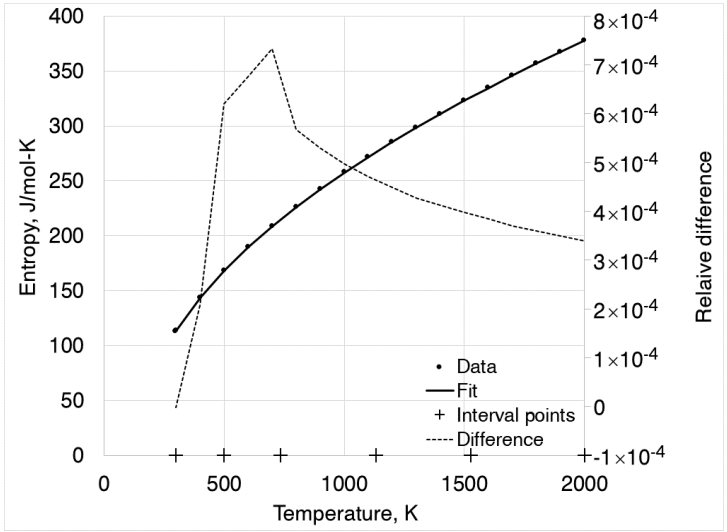


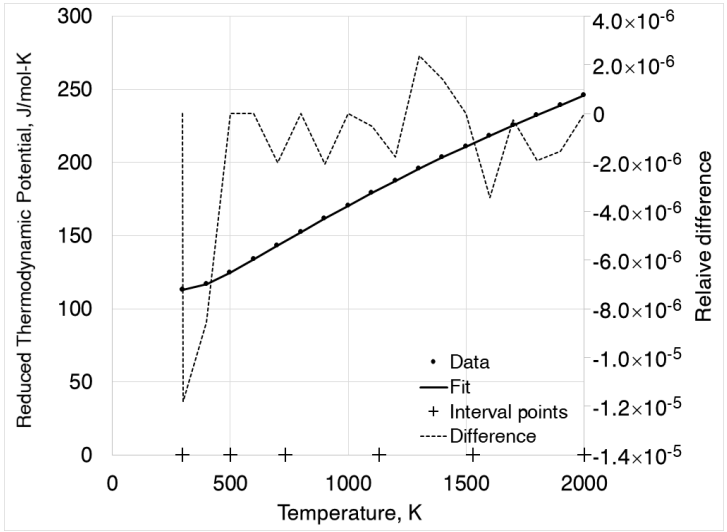
Potassium Peroxide (K_2O_2) – K2O2(cr)

MW = 110.1954 g/mol – $\Delta_f H^o$(298.15 K) = -495.804 kJ/mol – S^o(298.15 K) = 112.968 J/mol-K – p^o = 0.1 MPa



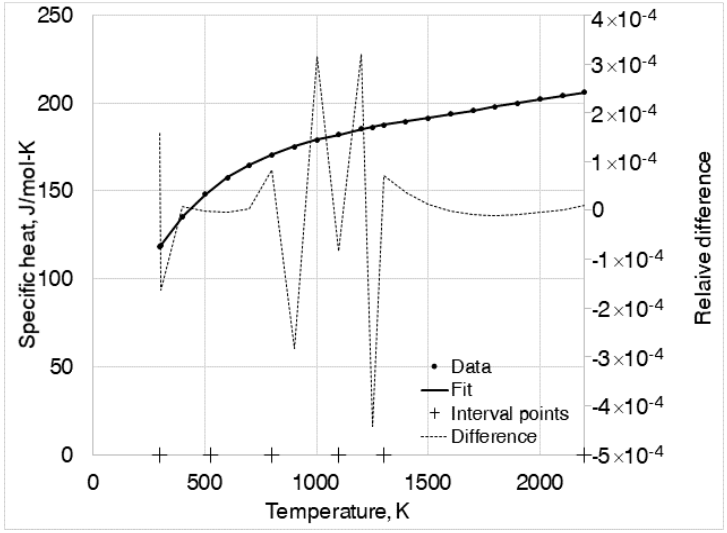


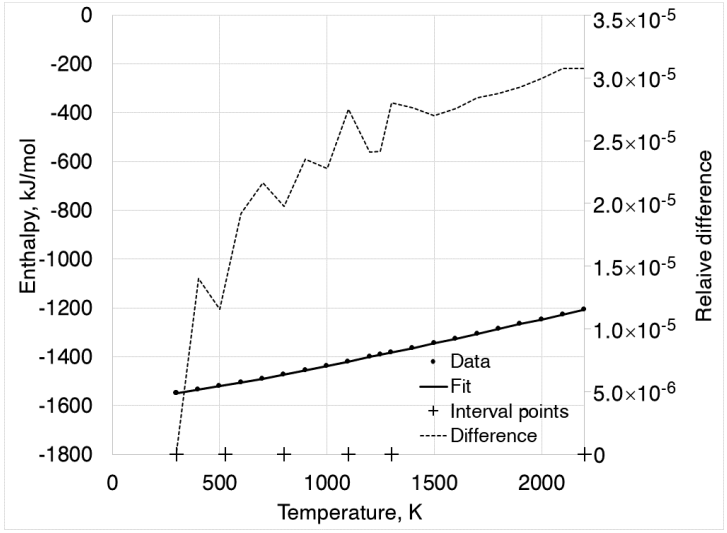


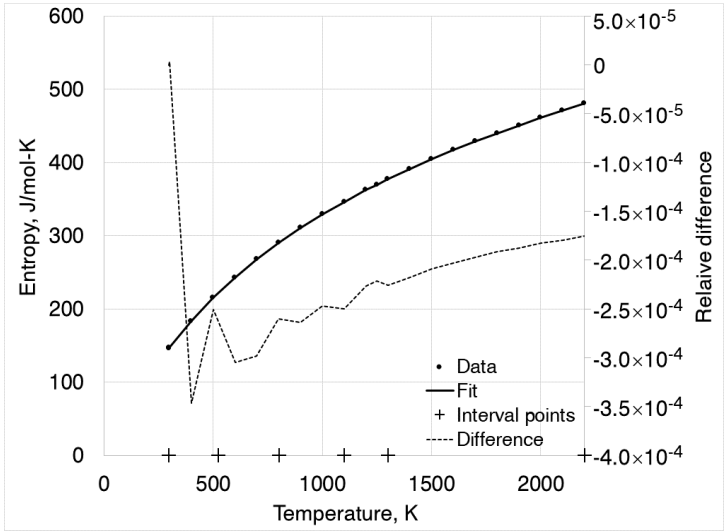


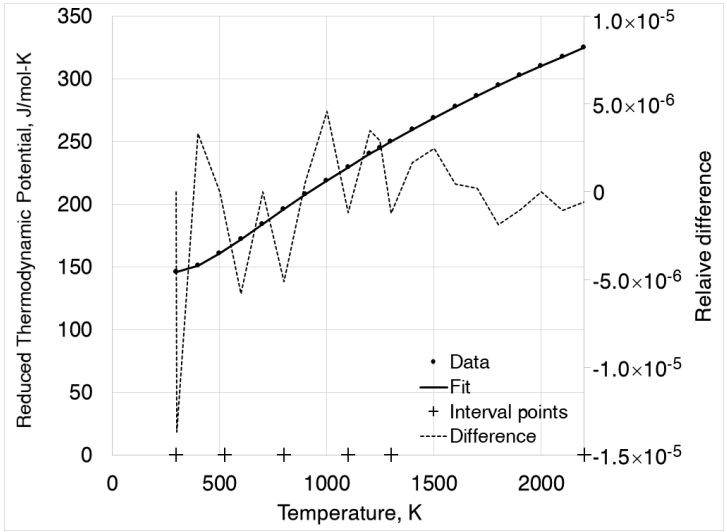
Potassium Silicate (K_2SiO_3) – K2O3Si1(cr)

MW = 154.2803 g/mol – $\Delta_f H^o$(298.15 K) = -1548.08 kJ/mol – S^o(298.15 K) = 146.147 J/mol-K – p^o = 0.1 MPa



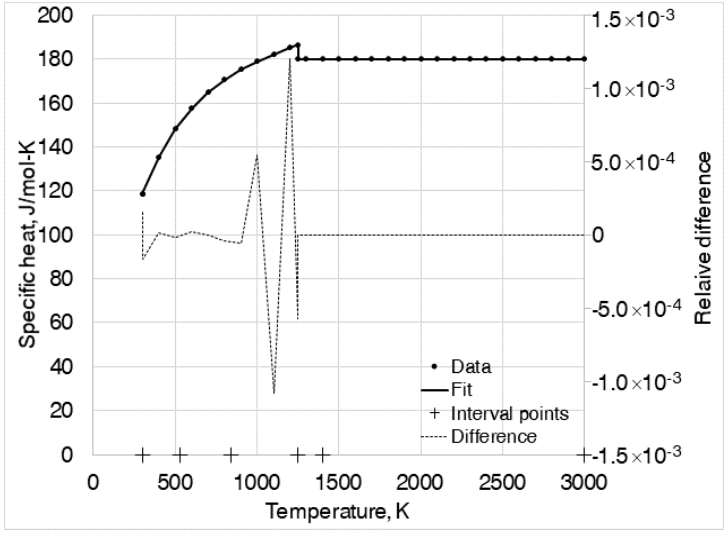


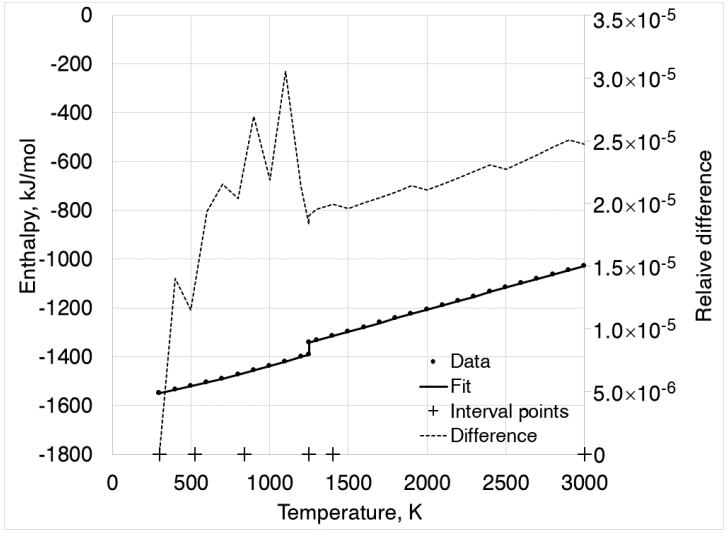


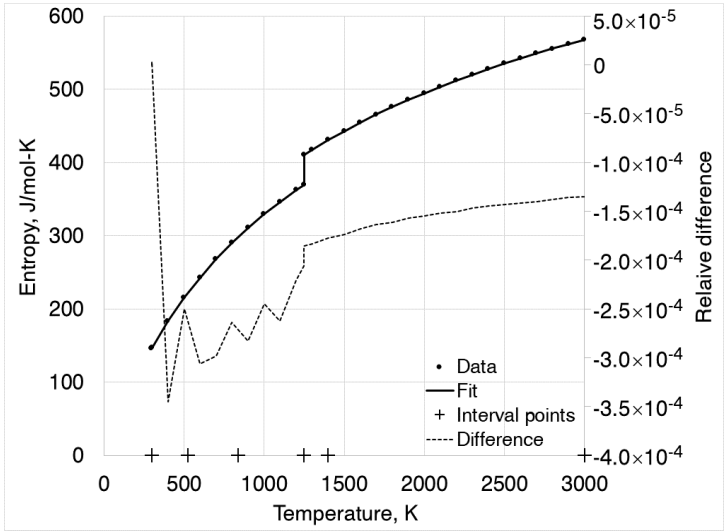


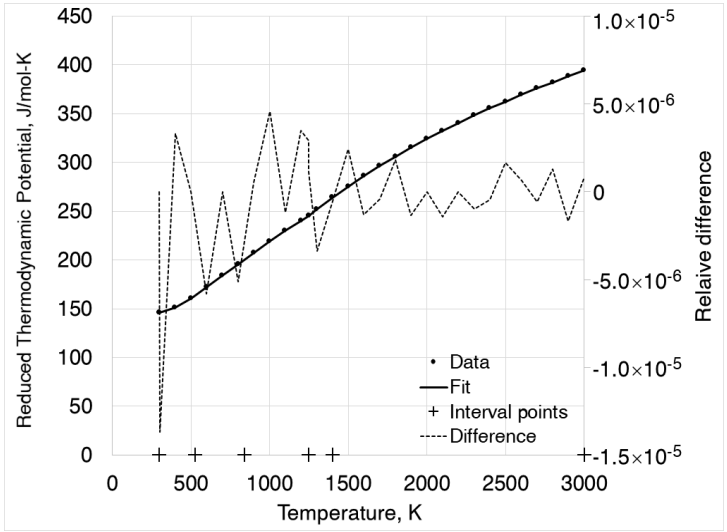
Potassium Silicate (K_2SiO_3) – K2O3Si1(cr,l)

MW = 154.2803 g/mol – $\Delta_f H^o$(298.15 K) = -1548.08 kJ/mol – S^o(298.15 K) = 146.147 J/mol-K – p^o = 0.1 MPa



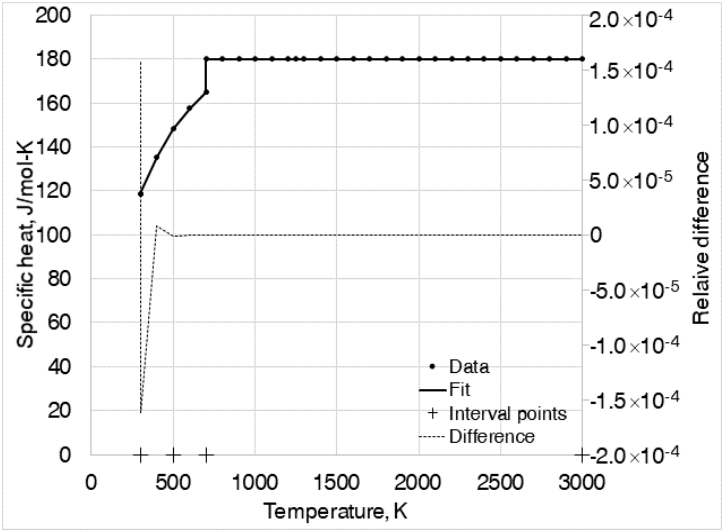


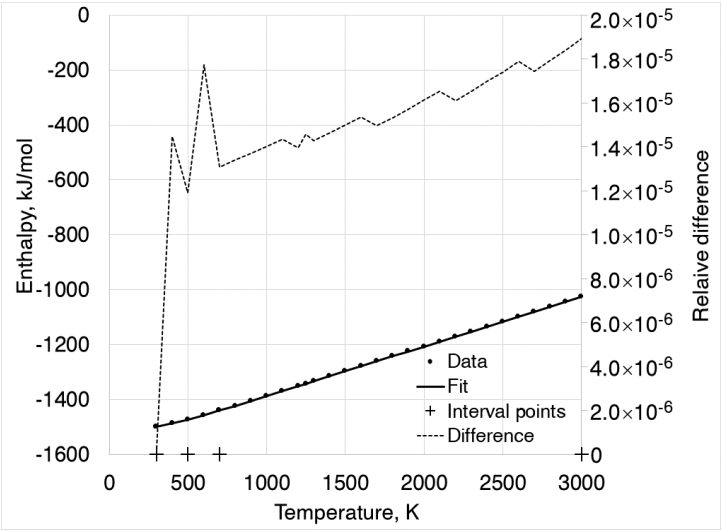


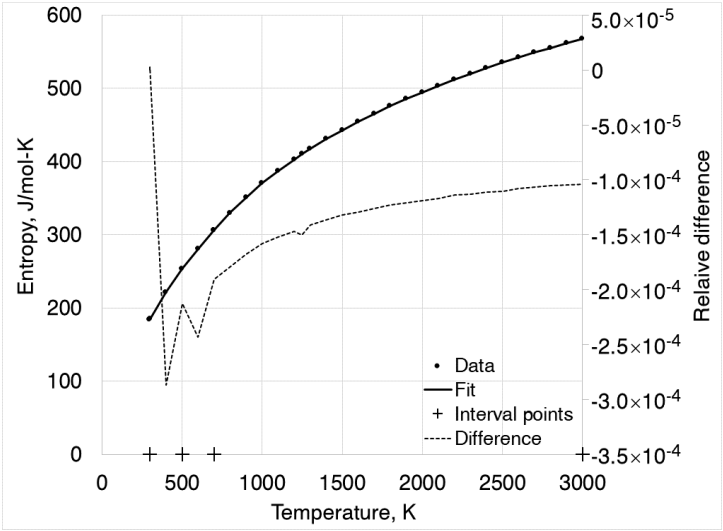


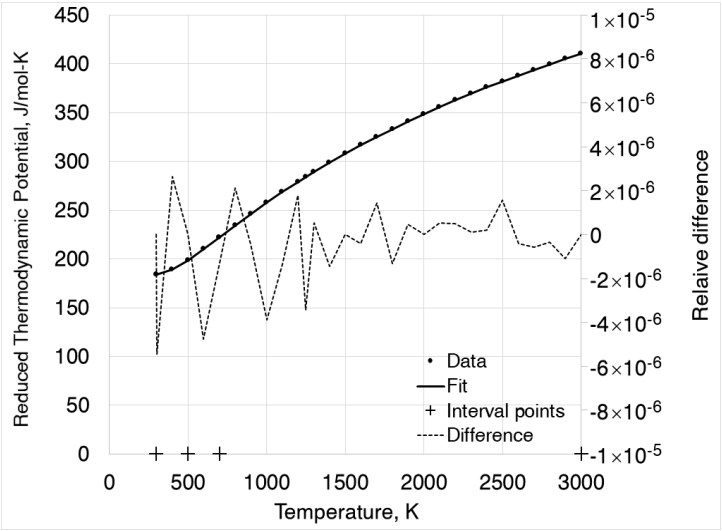
Potassium Silicate (K_2SiO_3) – K2O3Si1(l)

MW = 154.2803 g/mol – $\Delta_f H^o$(298.15 K) = -1499.291 kJ/mol – S^o(298.15 K) = 184.256 J/mol-K – p^o = 0.1 MPa



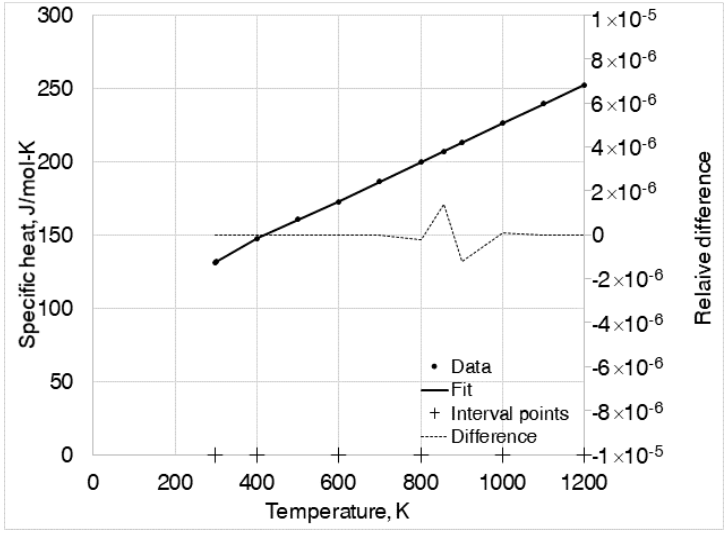


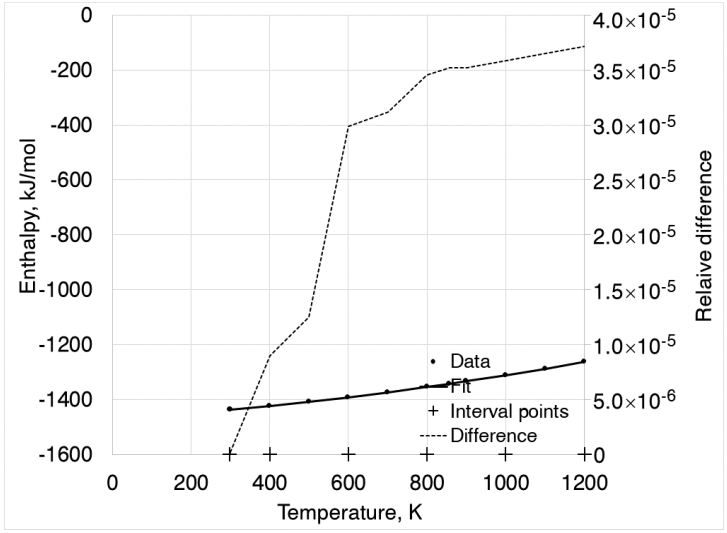


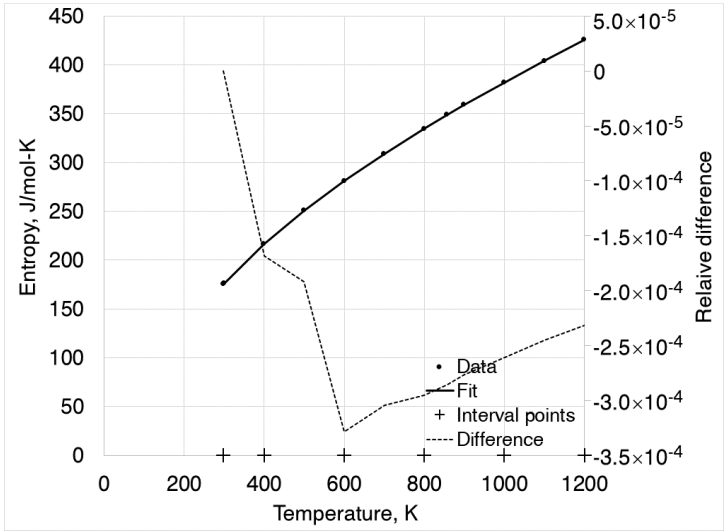


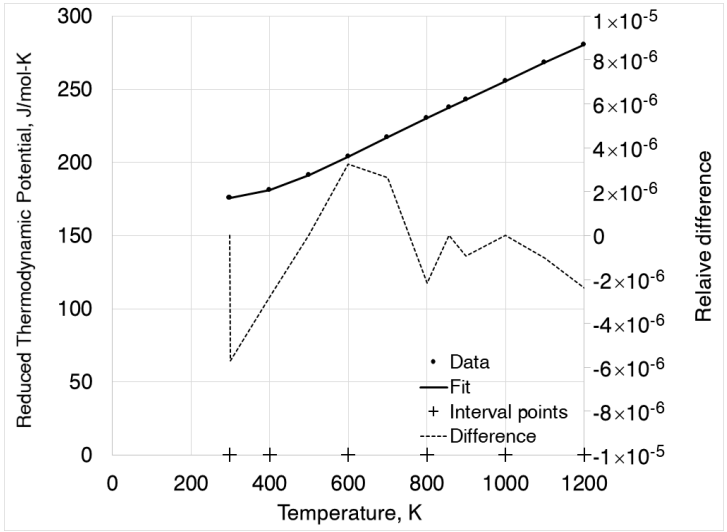
Potassium Sulfate, Alpha (K_2SO_4) – K2O4S1(cr)_01

MW = 174.25 g/mol – $\Delta_f H^o$(298.15 K) = -1437.706 kJ/mol – S^o(298.15 K) = 175.544 J/mol-K – p^o = 0.1 MPa



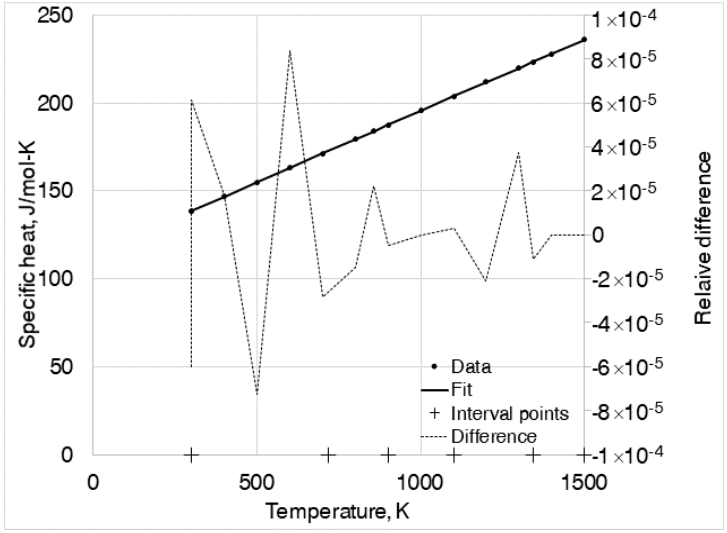


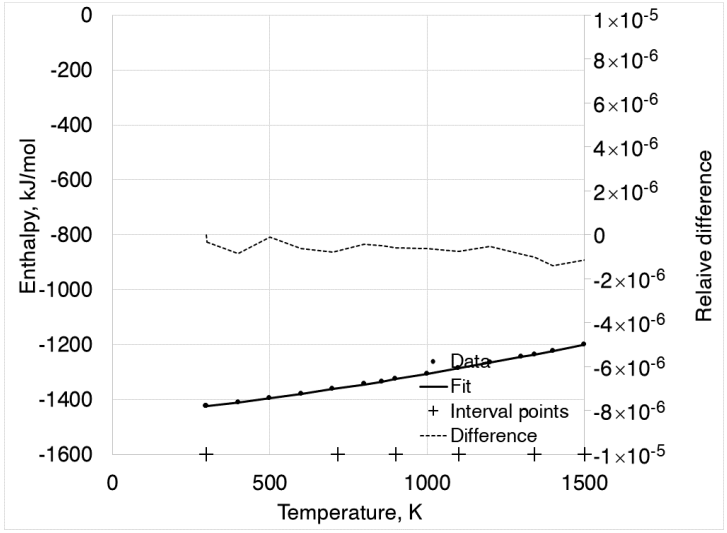


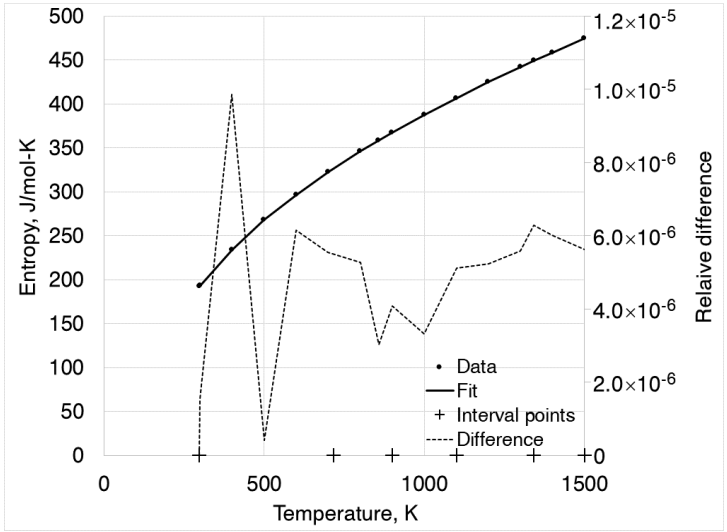


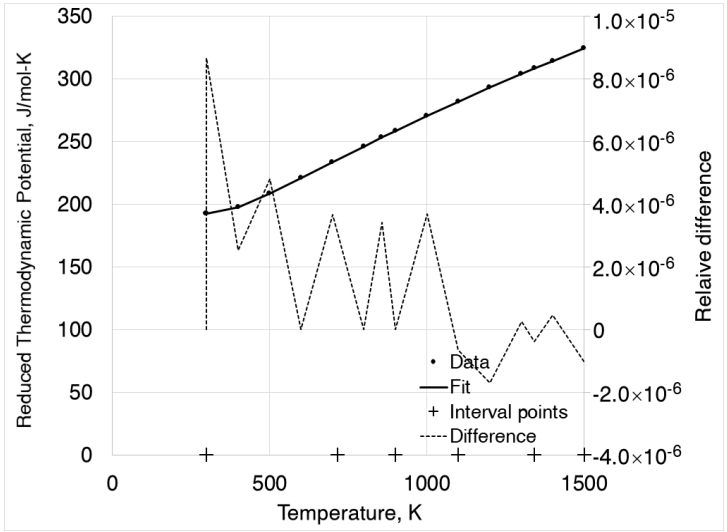
Potassium Sulfate, Beta (K_2SO_4) – K2O4S1(cr)_02

MW = 174.25 g/mol – $\Delta_f H^o$(298.15 K) = -1424.243 kJ/mol – S^o(298.15 K) = 192.279 J/mol-K – p^o = 0.1 MPa



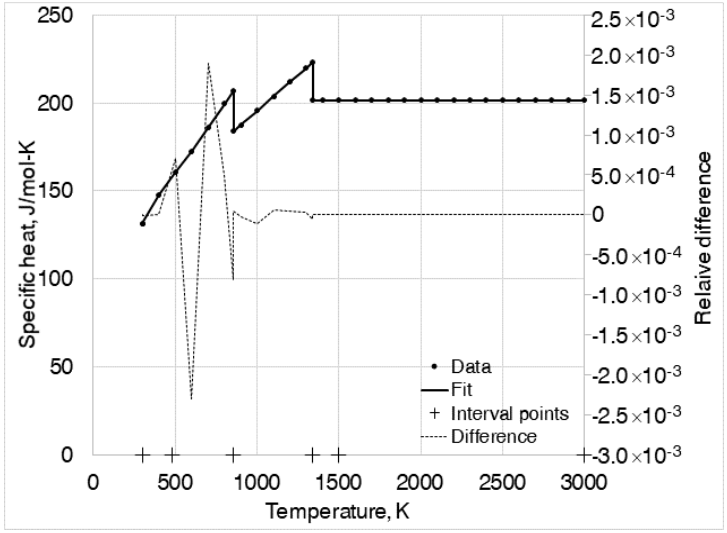


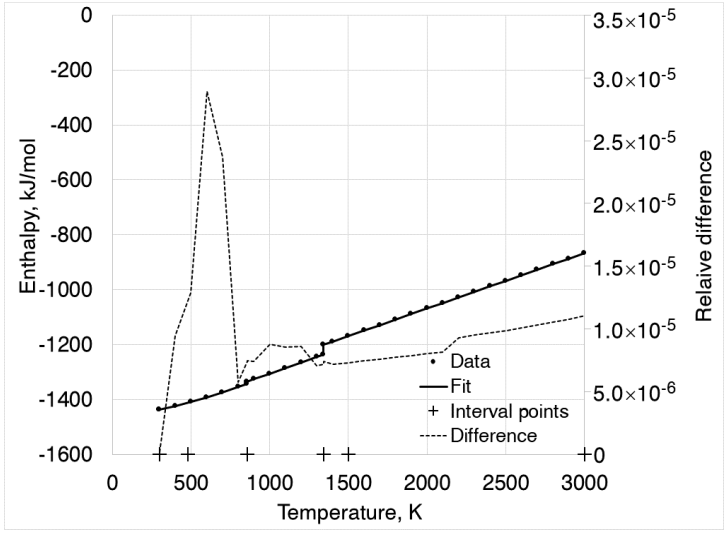


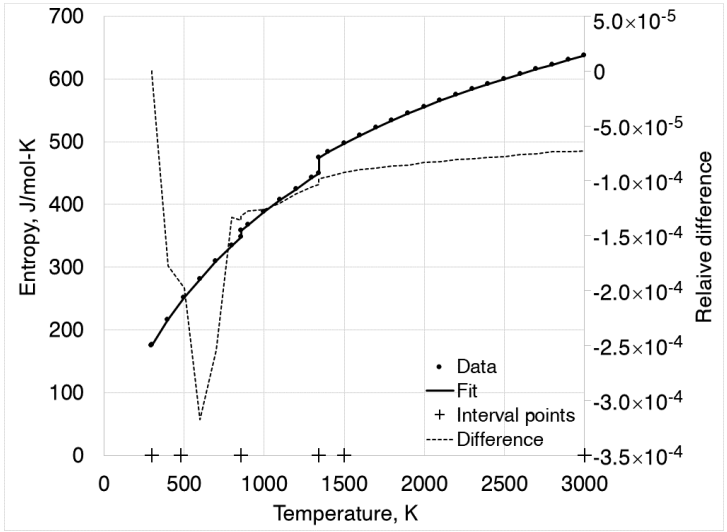


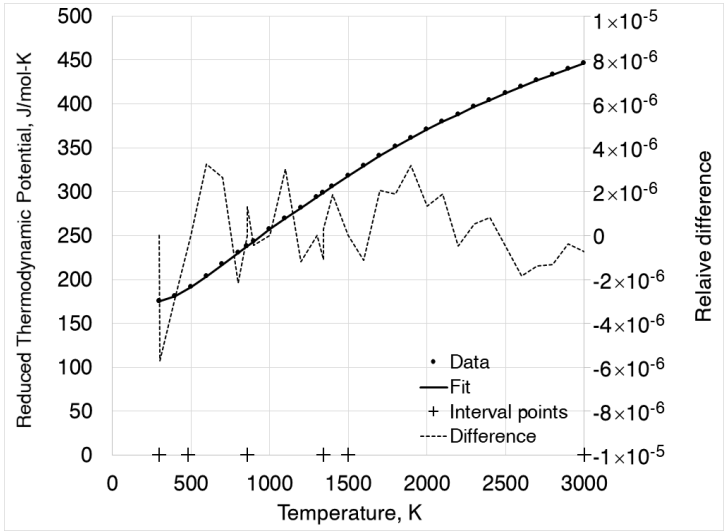
Potassium Sulfate (K_2SO_4) – K2O4S1(cr,l)

MW = 174.25 g/mol – $\Delta_f H^o$(298.15 K) = -1437.706 kJ/mol – S^o(298.15 K) = 175.544 J/mol-K – p^o = 0.1 MPa



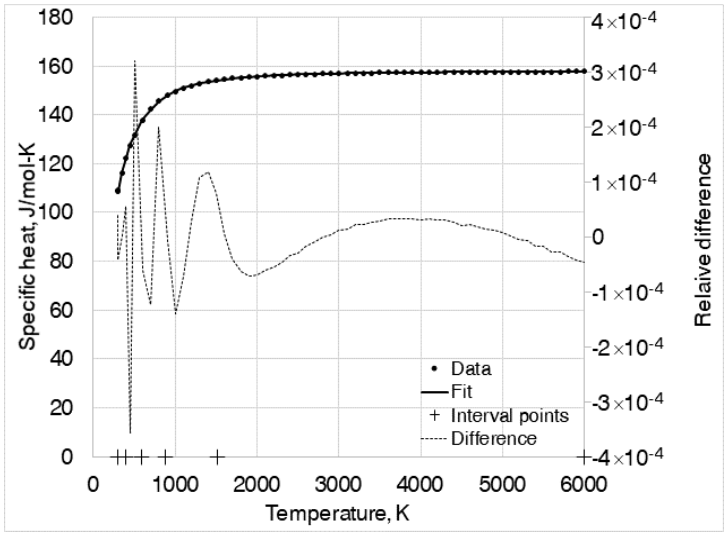


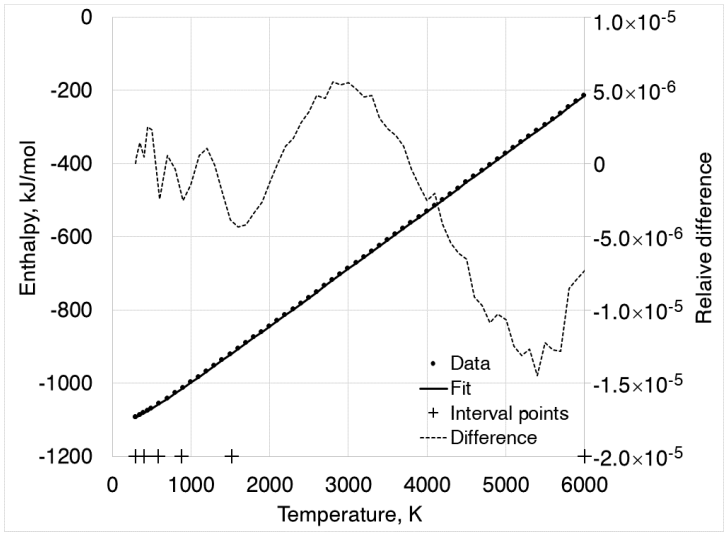


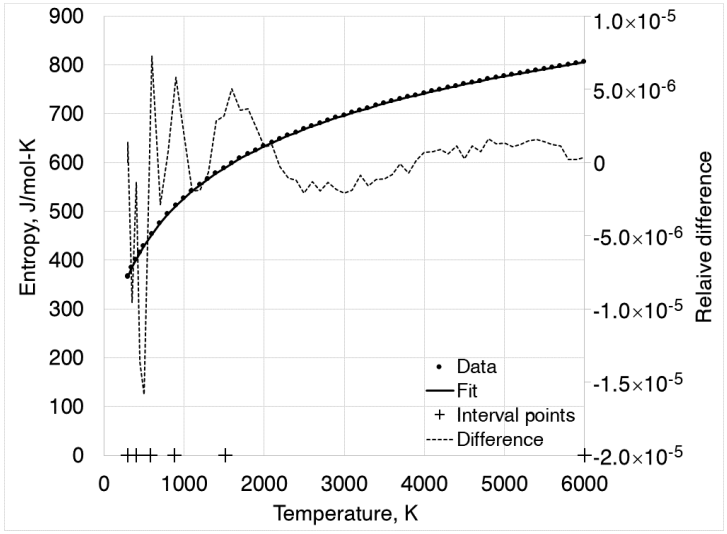


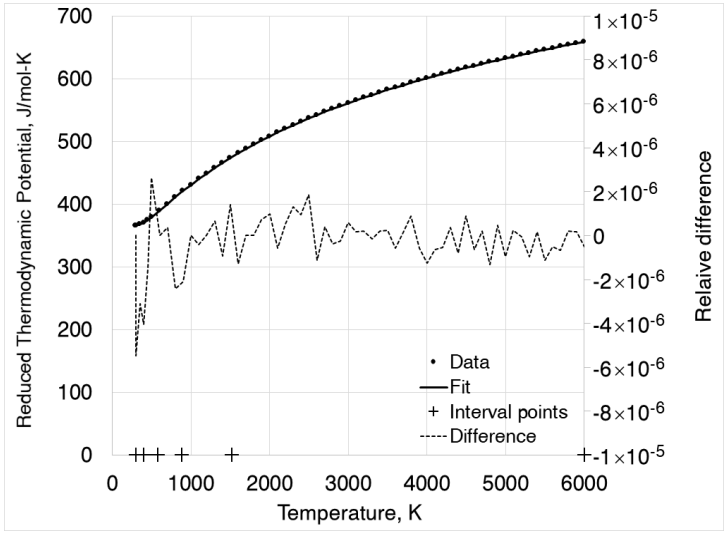
Potassium Sulfate (K_2SO_4) – K2O4S1(g)

MW = 174.25 g/mol – $\Delta_f H^o$(298.15 K) = -1094.116 kJ/mol – S^o(298.15 K) = 366.152 J/mol-K – p^o = 0.1 MPa



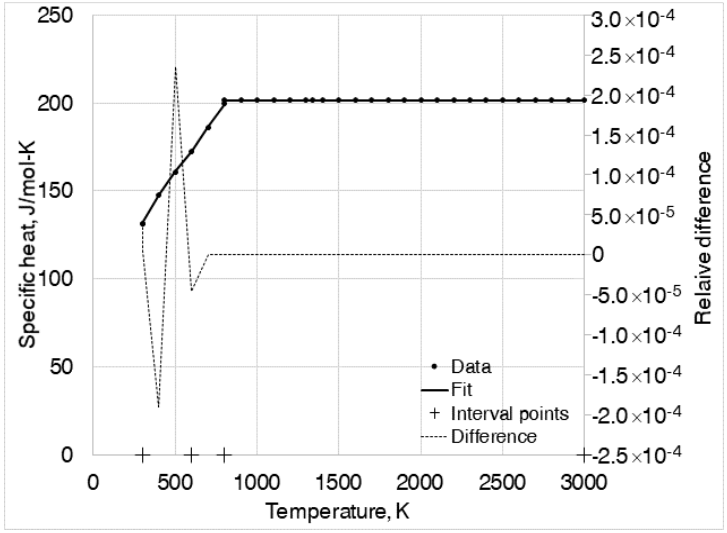


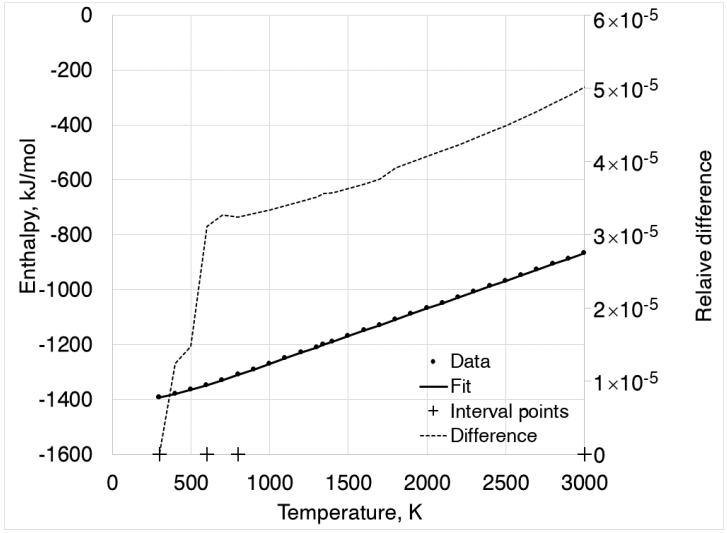


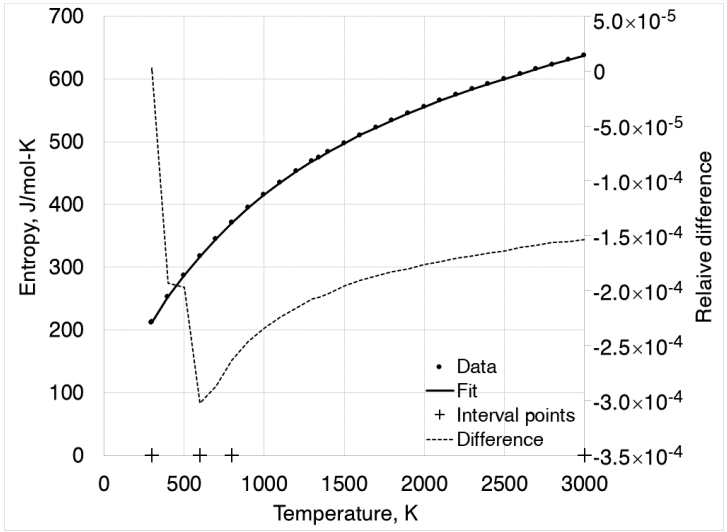


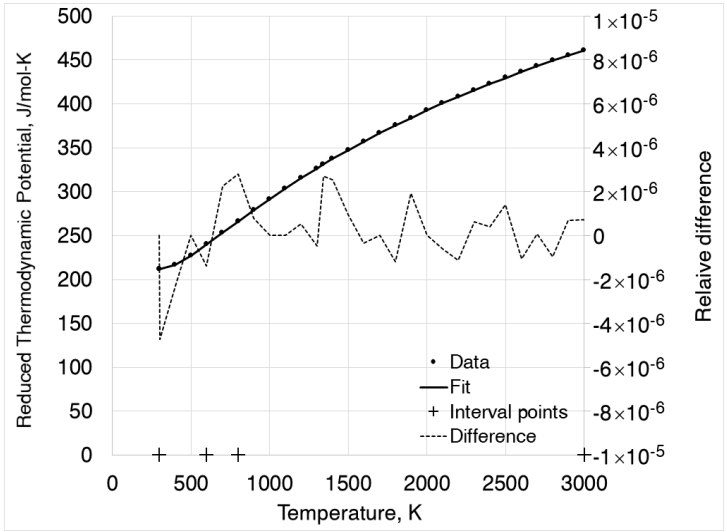
Potassium Sulfate (K_2SO_4) – K2O4S1(l)

MW = 174.25 g/mol – $\Delta_f H^o$(298.15 K) = -1393.668 kJ/mol – S^o(298.15 K) = 211.505 J/mol-K – p^o = 0.1 MPa



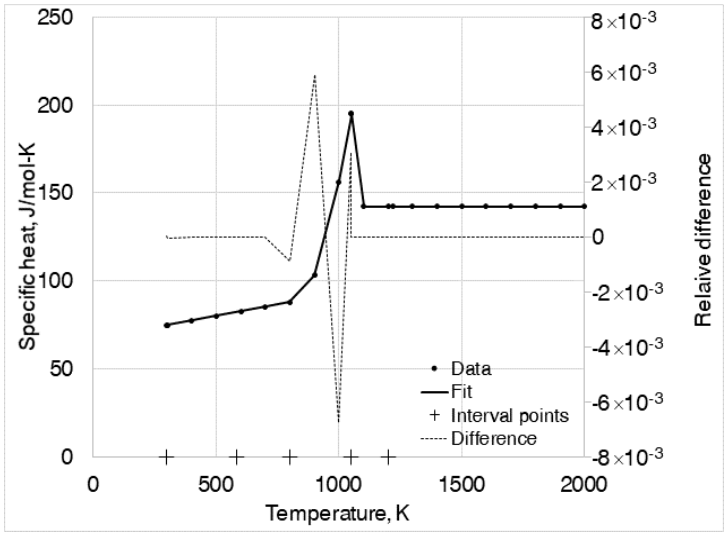


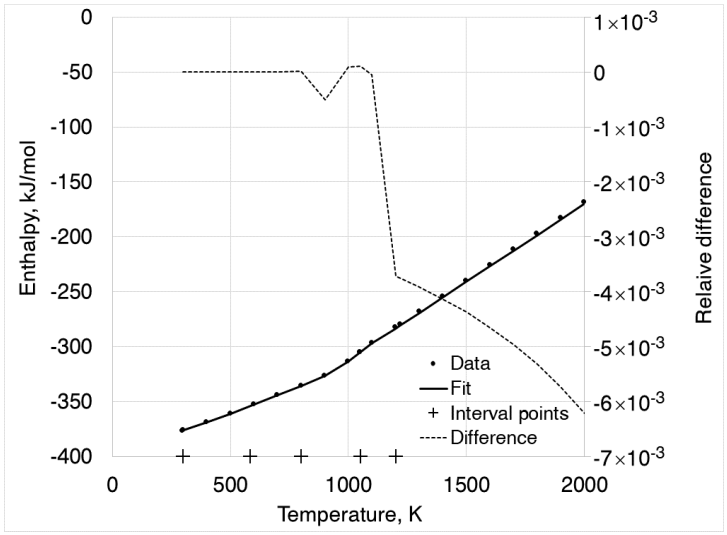


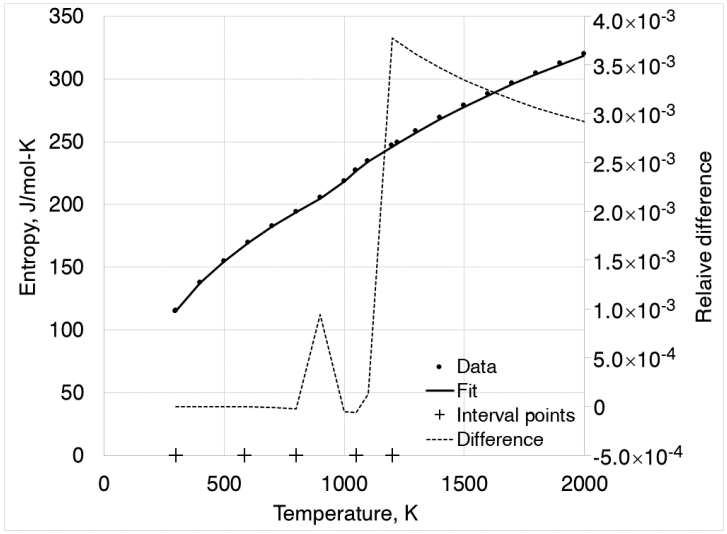


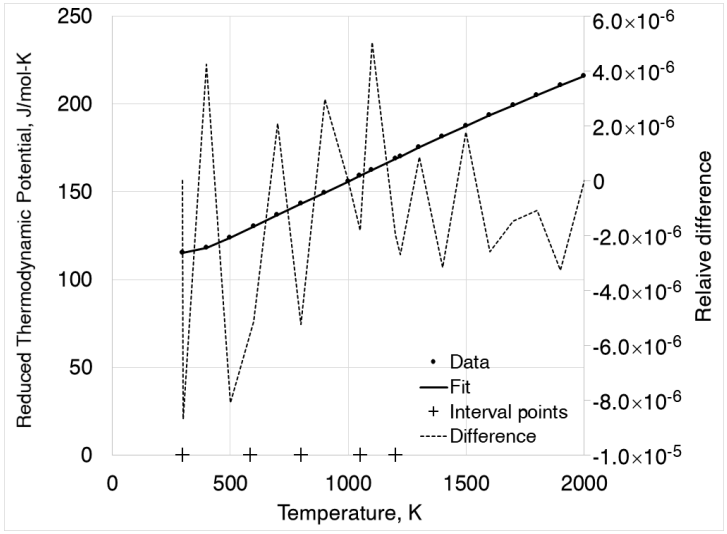
Potassium Sulfide (K_2S) – K2S1(cr)

MW = 110.26 g/mol – $\Delta_f H^o$(298.15 K) = -376.56 kJ/mol – S^o(298.15 K) = 115.06 J/mol-K – p^o = 0.1 MPa



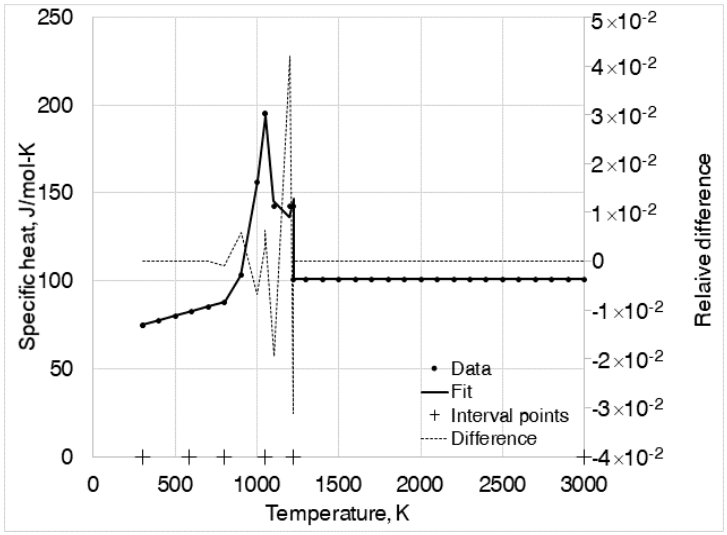


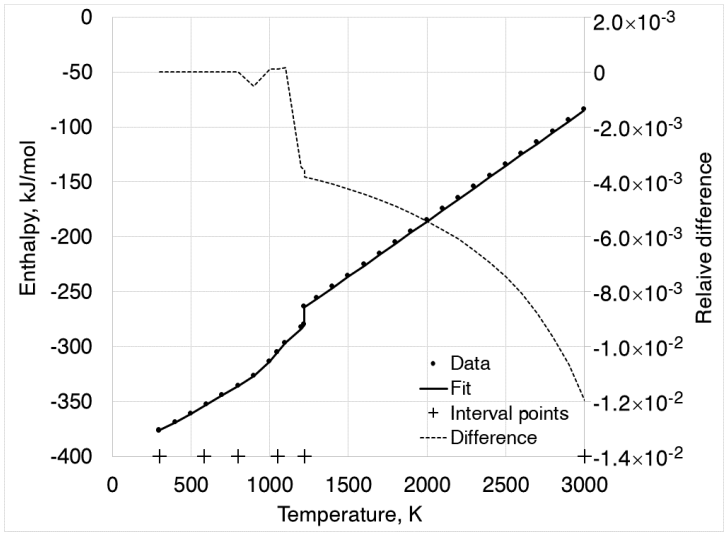


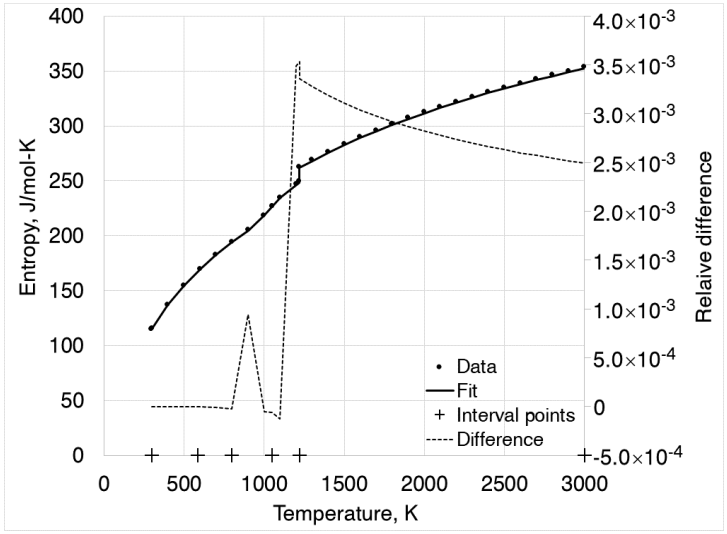


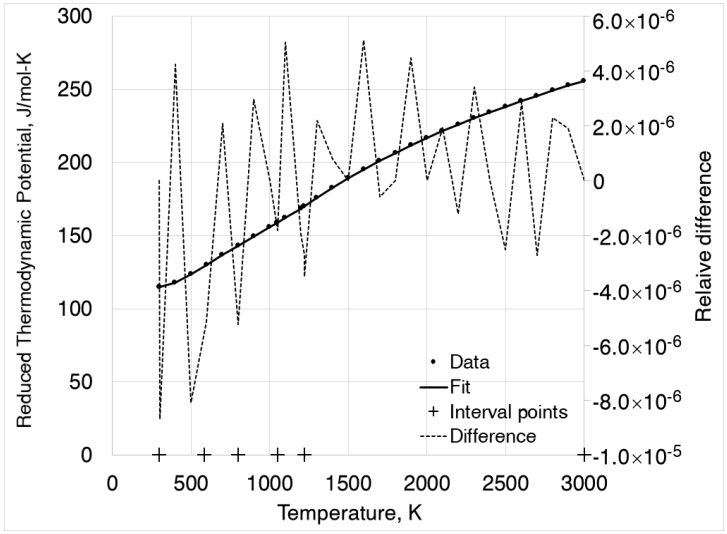
Potassium Sulfide (K_2S) – K2S1(cr,l)

MW = 110.26 g/mol – $\Delta_f H^o$(298.15 K) = -376.56 kJ/mol – S^o(298.15 K) = 115.06 J/mol-K – p^o = 0.1 MPa



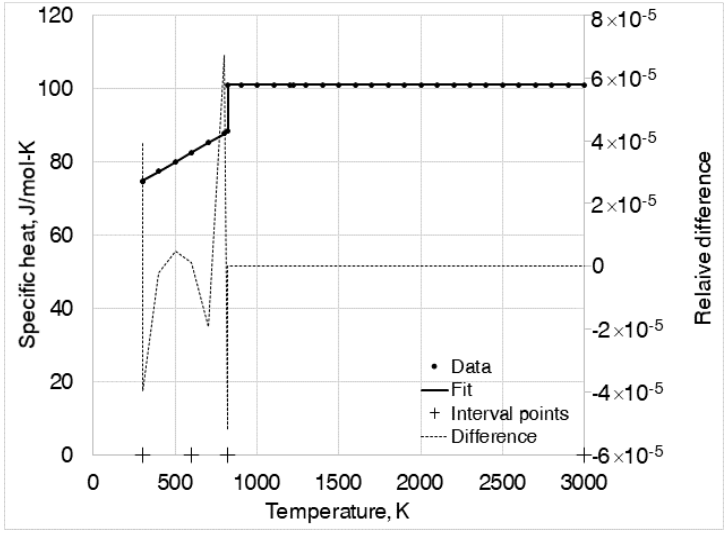


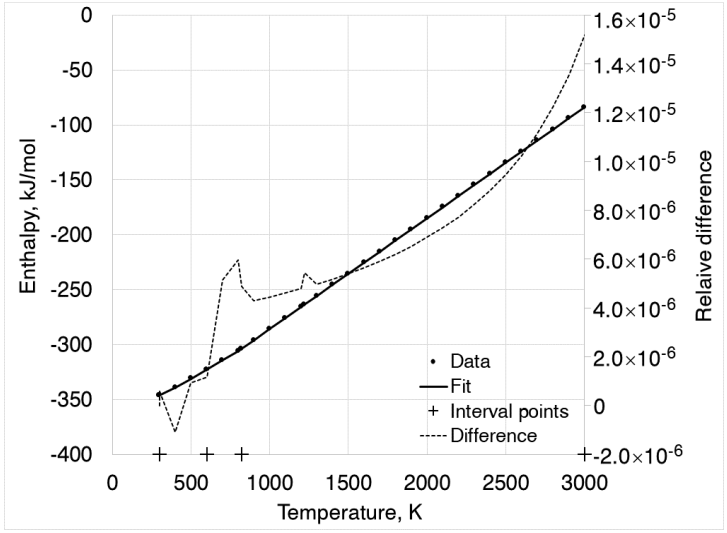


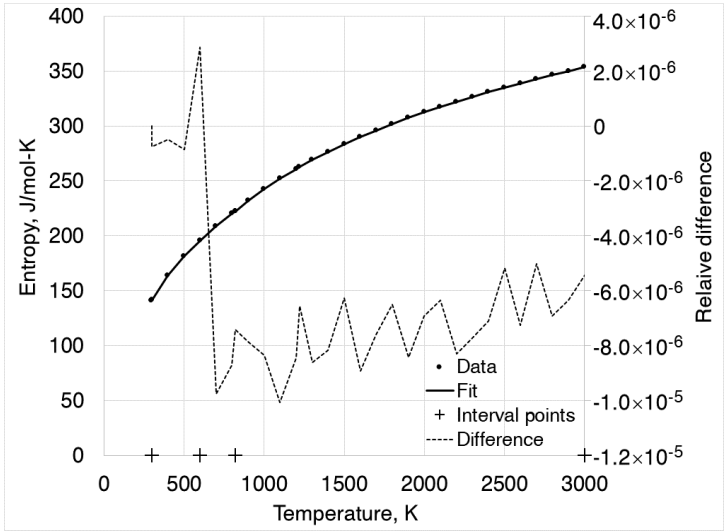


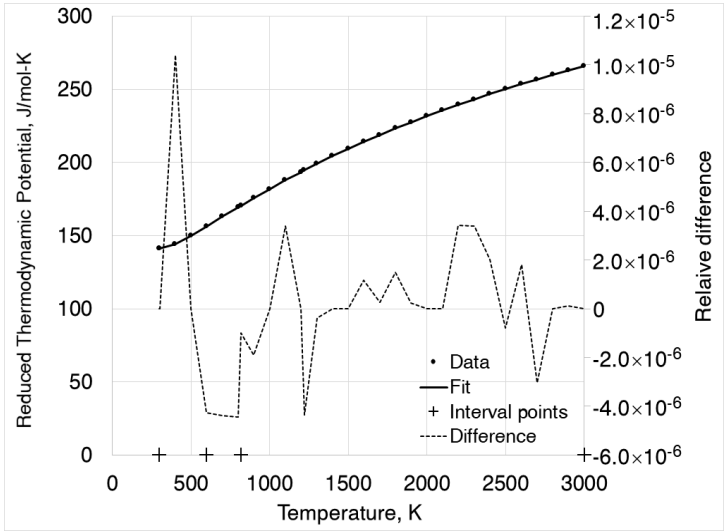
Potassium Sulfide (K_2S) – K2S1(l)

MW = 110.26 g/mol – $\Delta_f H^o$(298.15 K) = -346.506 kJ/mol – S^o(298.15 K) = 141.172 J/mol-K – p^o = 0.1 MPa



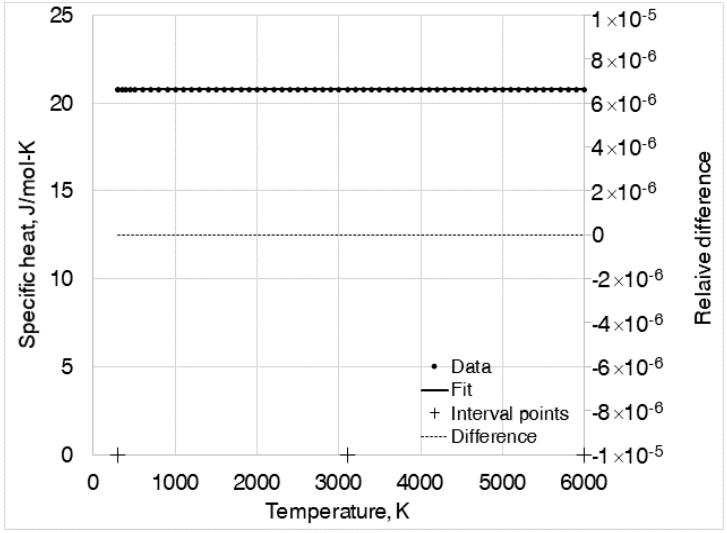


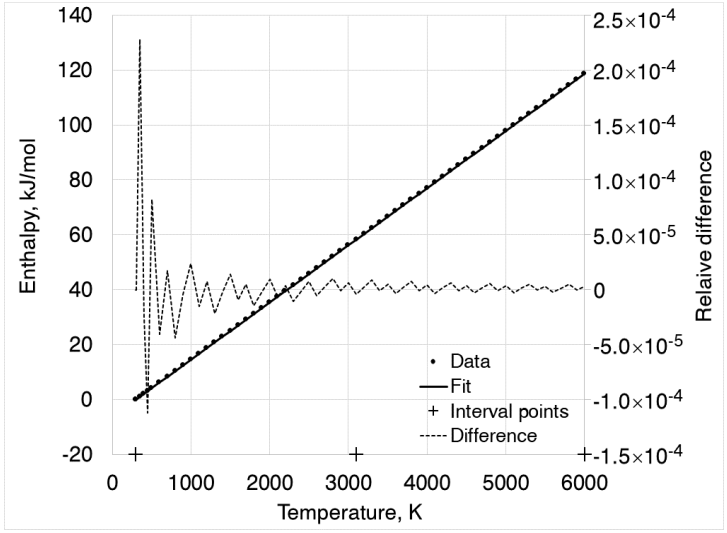


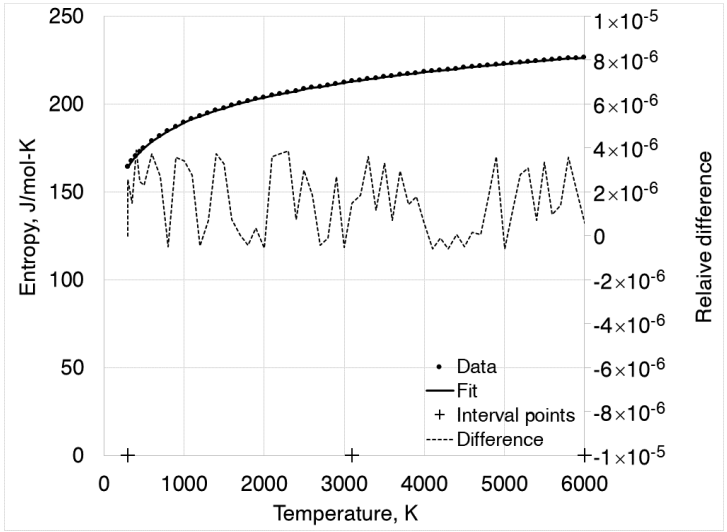


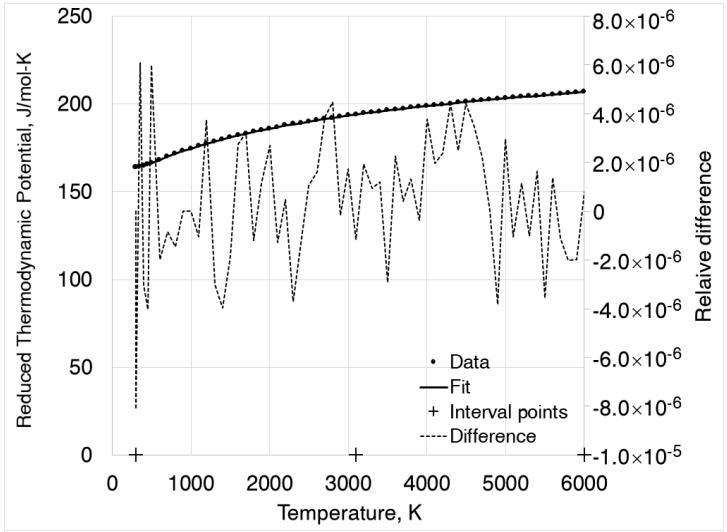
Krypton (Kr) – Kr1(ref)

MW = 83.8 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 164.084 J/mol-K – p^o = 0.1 MPa



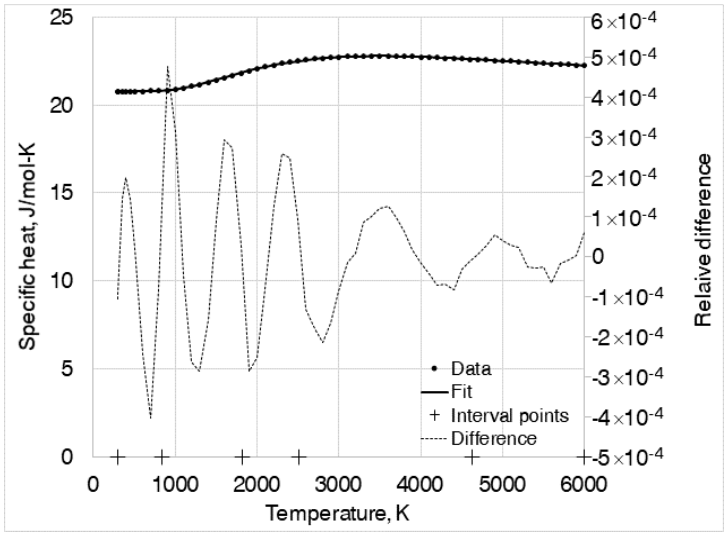


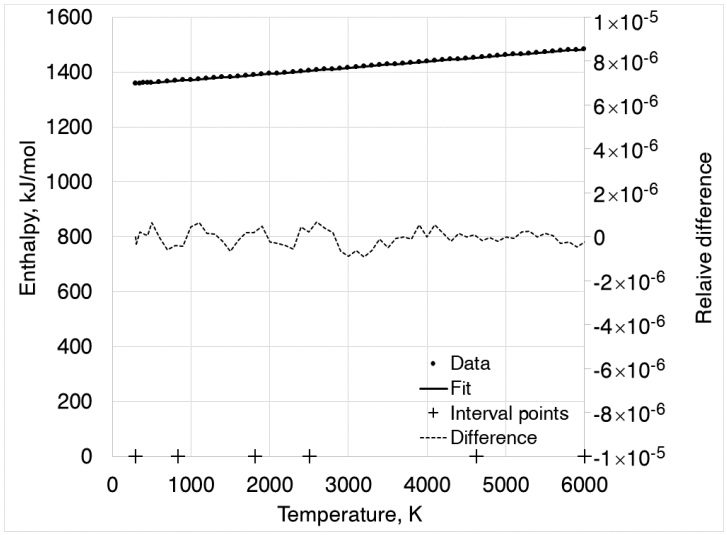


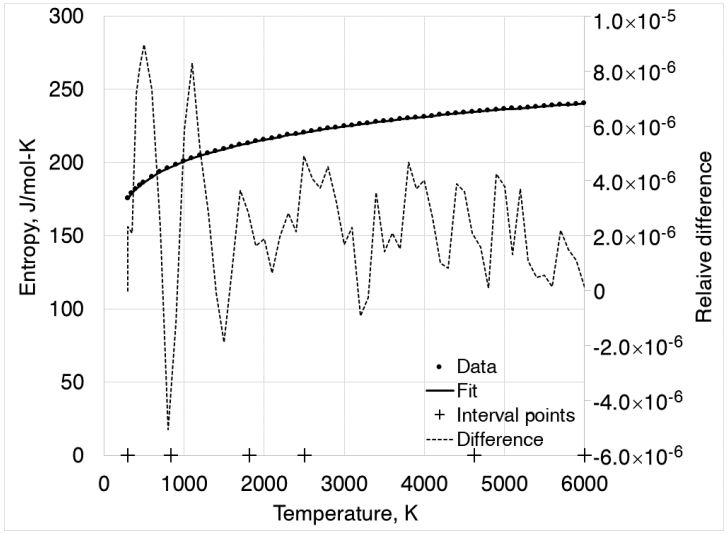


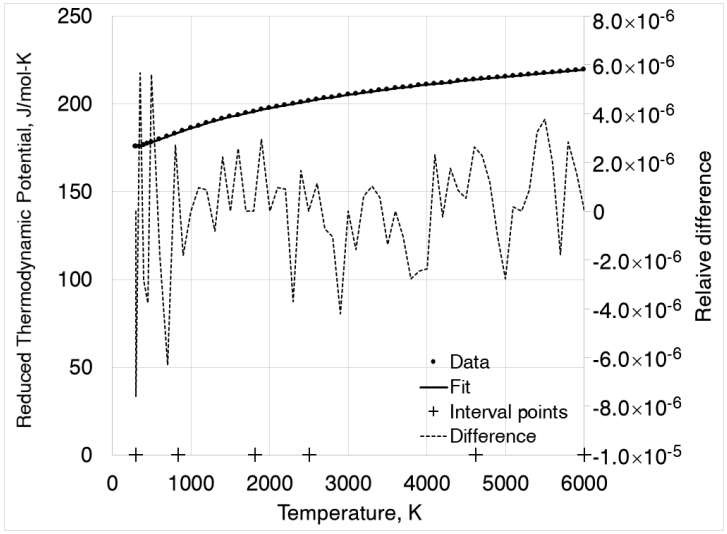
Krypton, Ion (Kr^+) – Kr1+(g)

MW = 83.8 g/mol – $\Delta_f H^o$(298.15 K) = 1356.955 kJ/mol – S^o(298.15 K) = 175.61 J/mol-K – p^o = 0.1 MPa



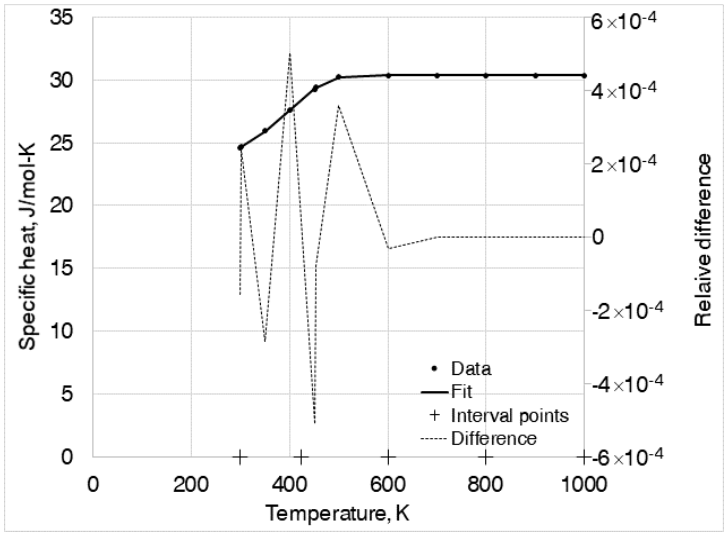


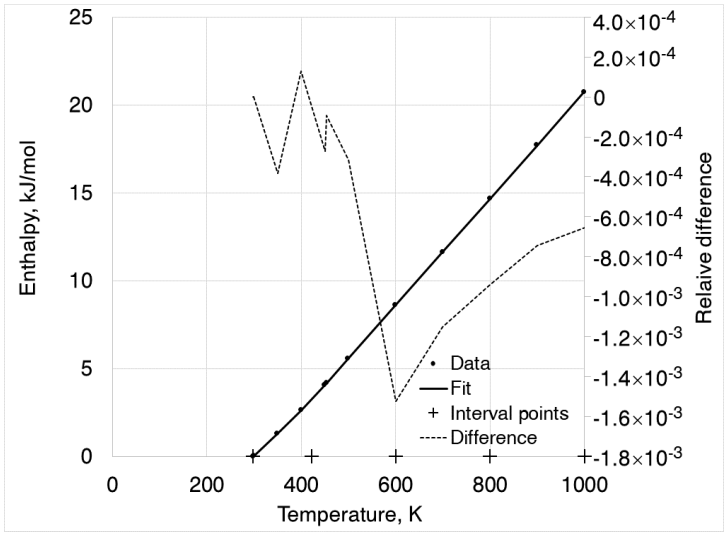


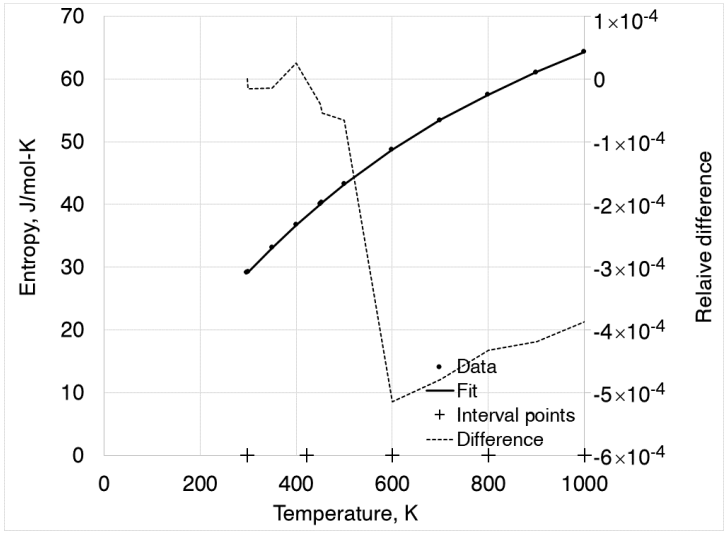


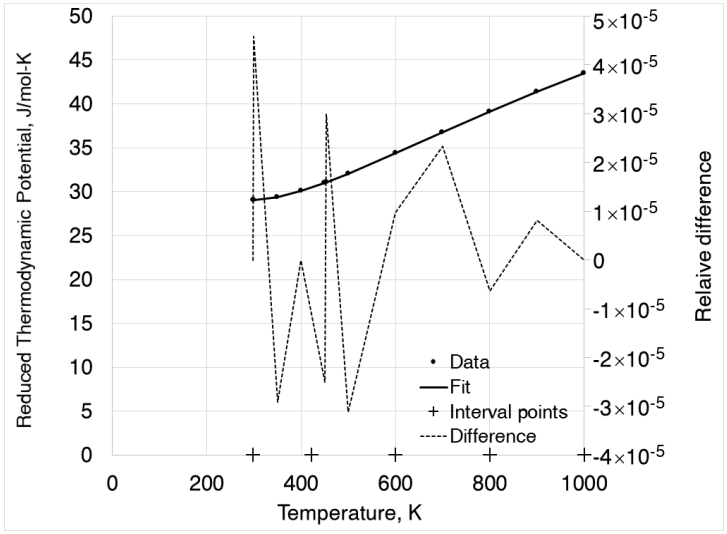
Lithium (Li) – Li1(cr)

MW = 6.941 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 29.085 J/mol-K – p^o = 0.1 MPa



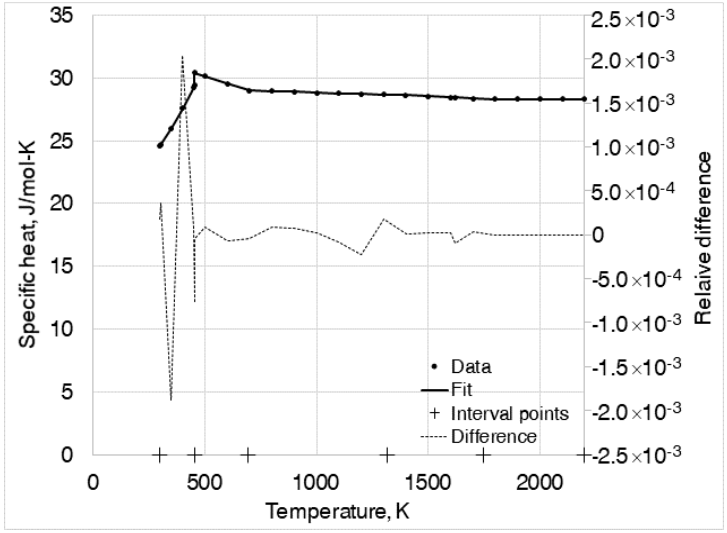


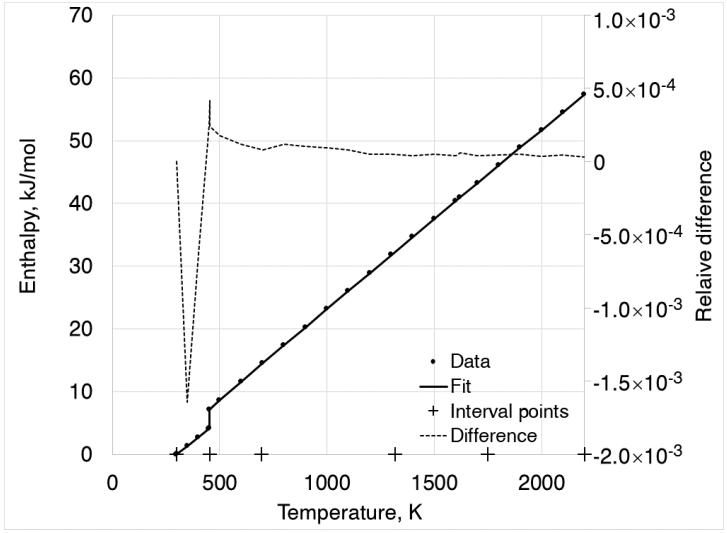


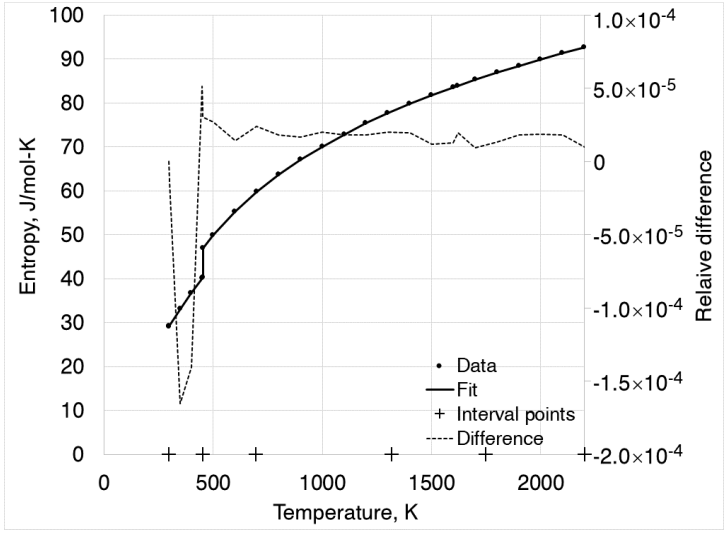


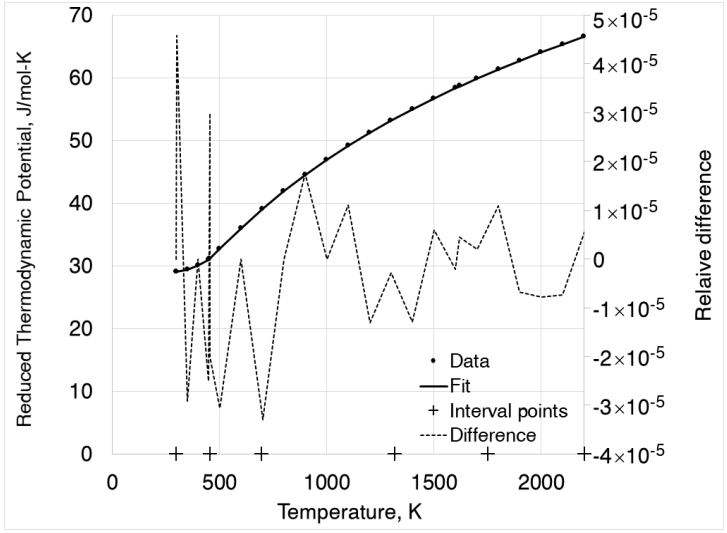
Lithium (Li) – Li1(cr,l)

MW = 6.941 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 29.085 J/mol-K – p^o = 0.1 MPa



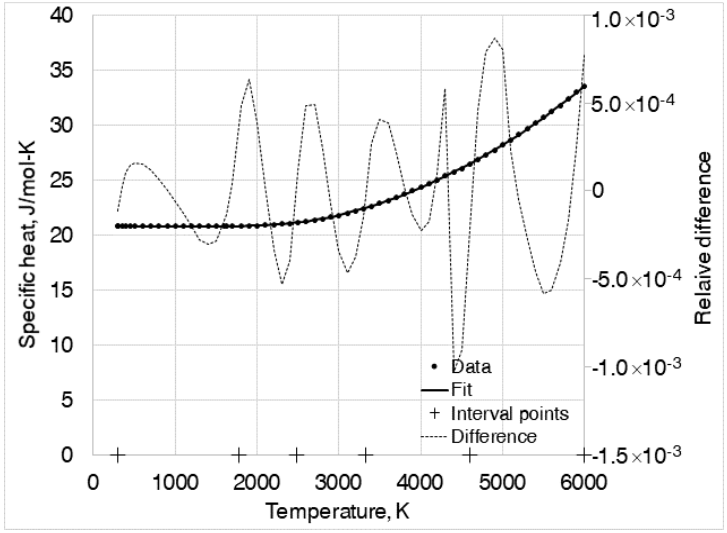


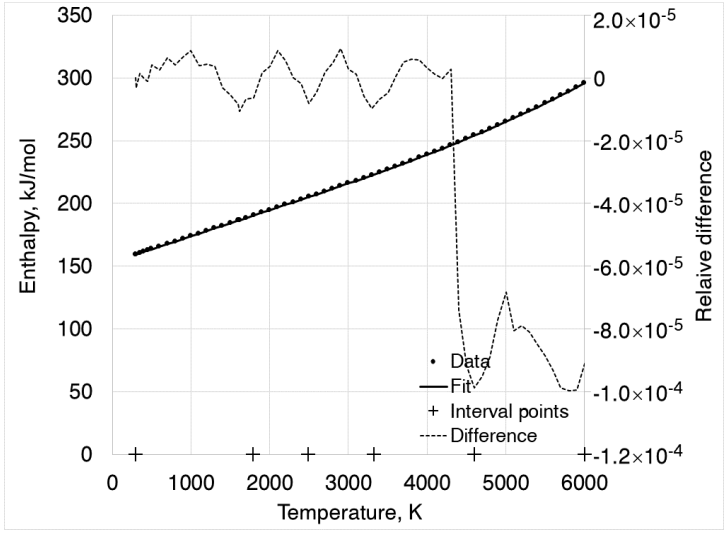


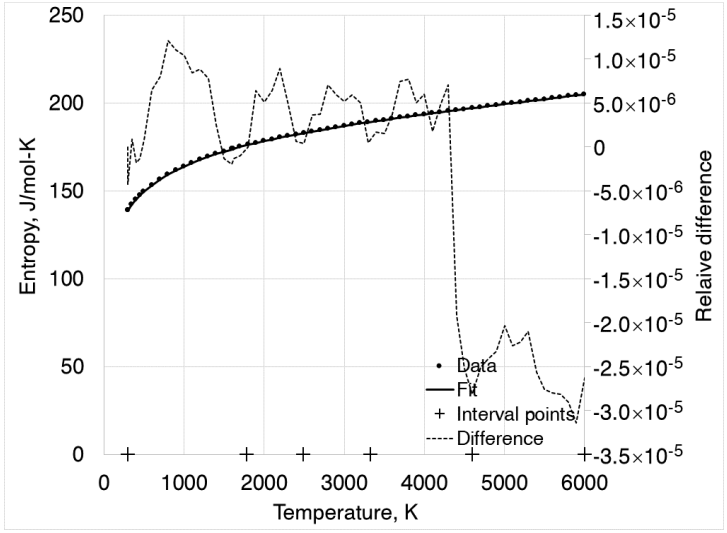


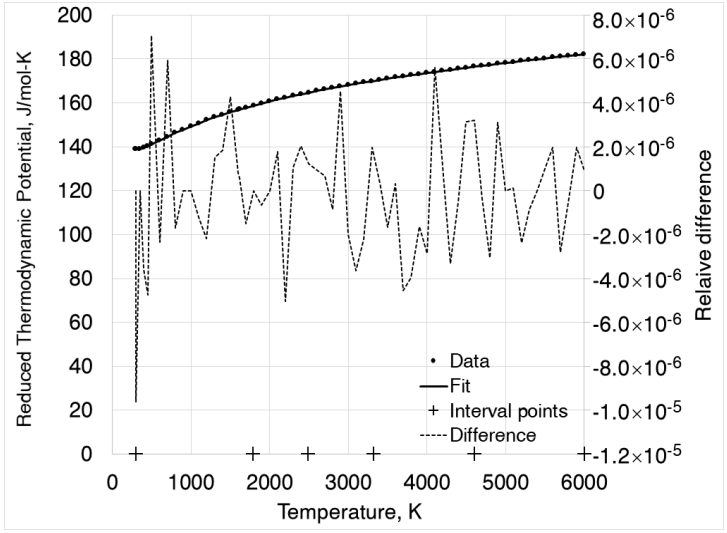
Lithium (Li) – Li1(g)

MW = 6.941 g/mol – $\Delta_f H^o$(298.15 K) = 159.3 kJ/mol – S^o(298.15 K) = 138.781 J/mol-K – p^o = 0.1 MPa



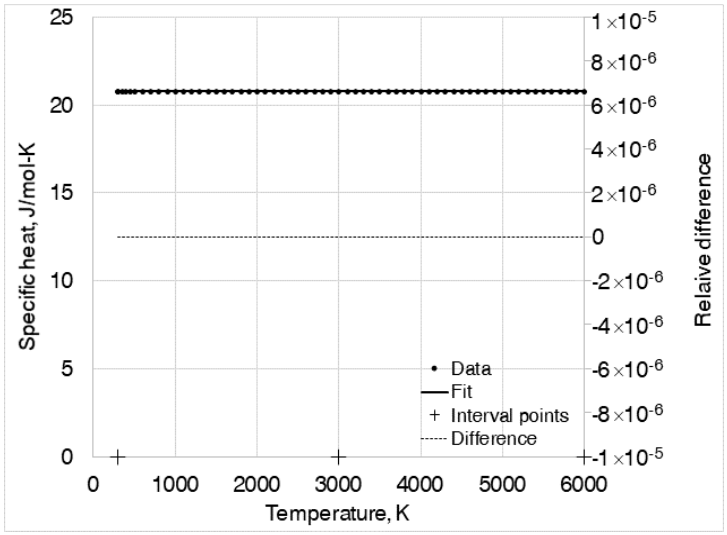


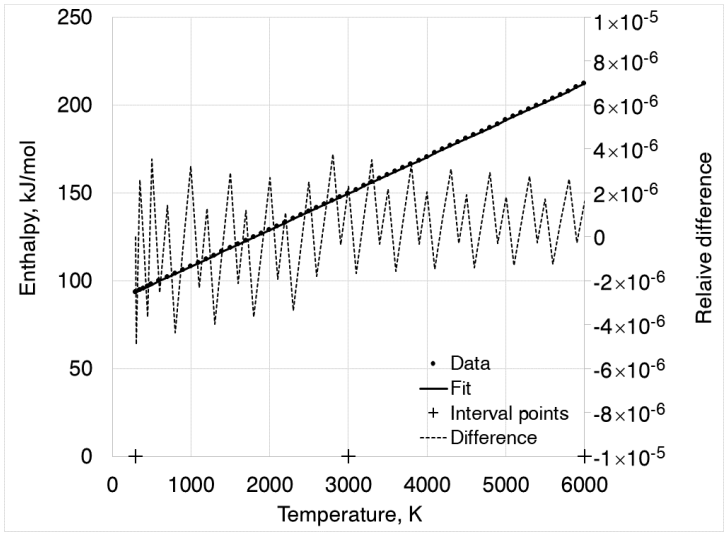


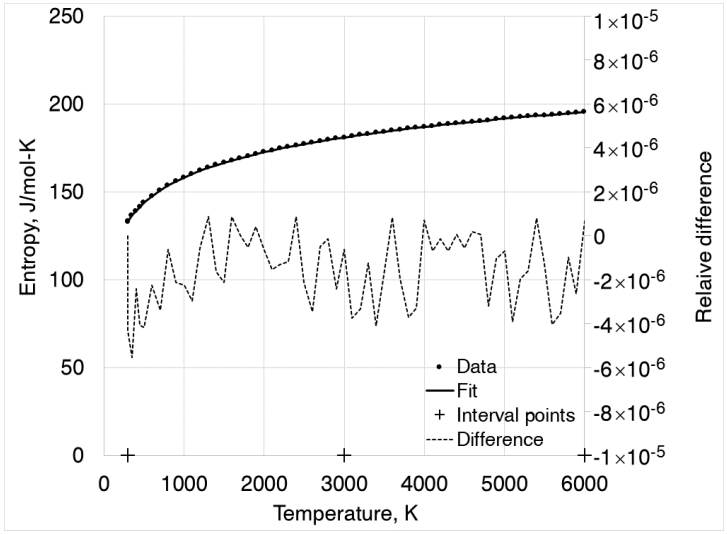


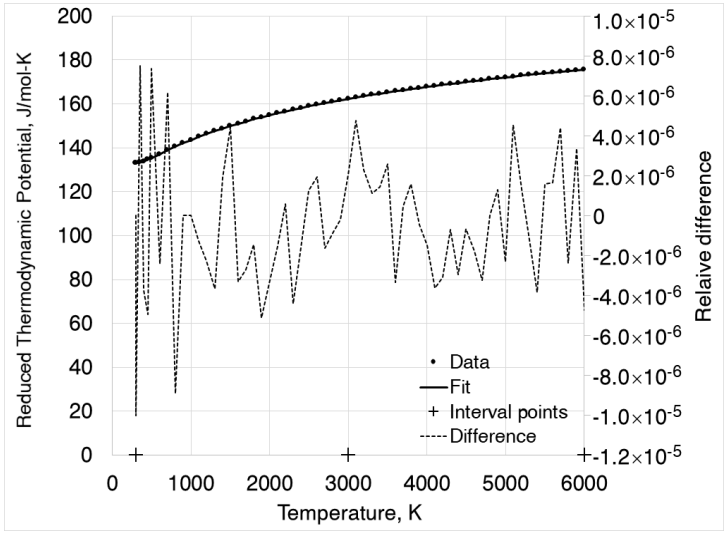
Lithium, Ion (Li$^-$) – Li1-(g)

MW = 6.942 g/mol – $\Delta_f H^o$(298.15 K) = 93.475 kJ/mol – S^o(298.15 K) = 133.019 J/mol-K – p^o = 0.1 MPa



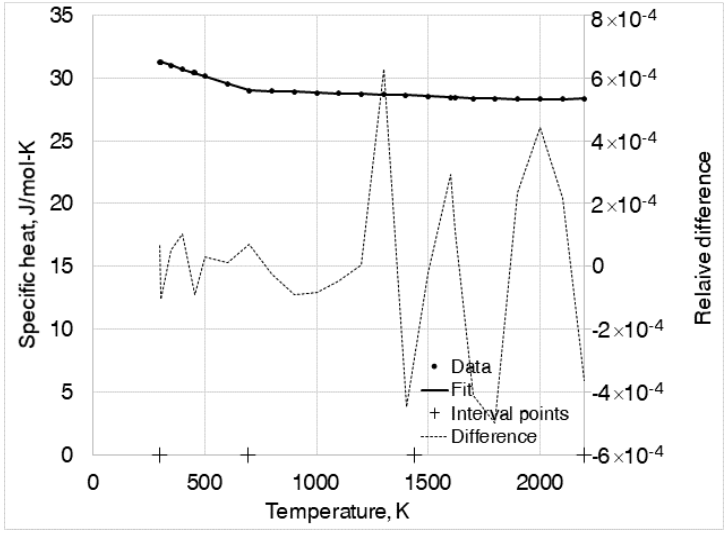


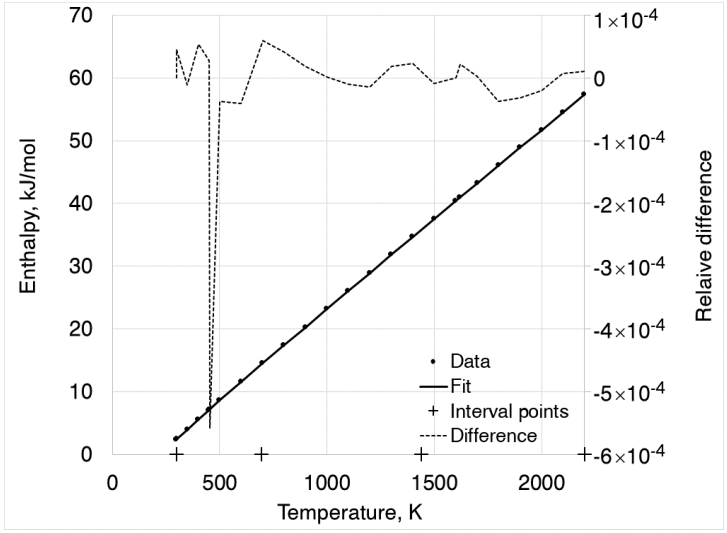


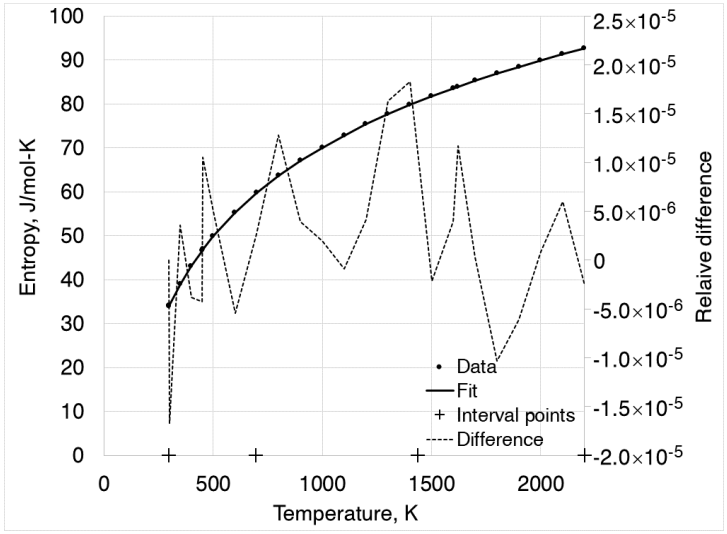


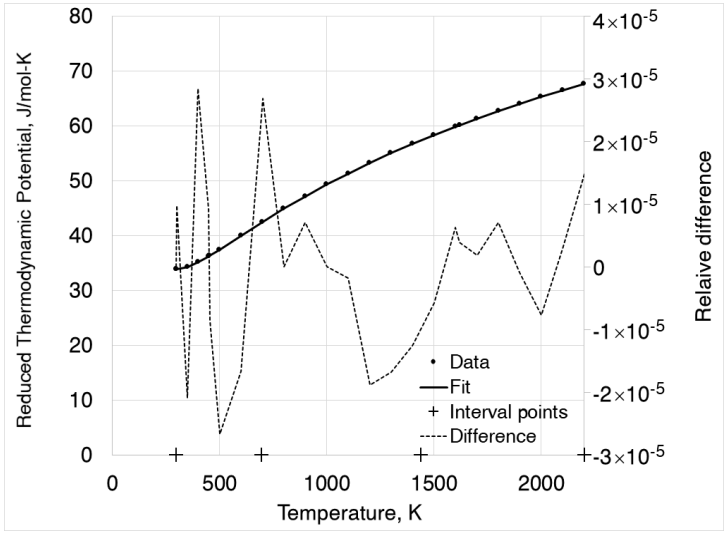
Lithium (Li) – Li1(l)

MW = 6.941 g/mol – $\Delta_f H^o$(298.15 K) = 2.38 kJ/mol – S^o(298.15 K) = 33.938 J/mol-K – p^o = 0.1 MPa



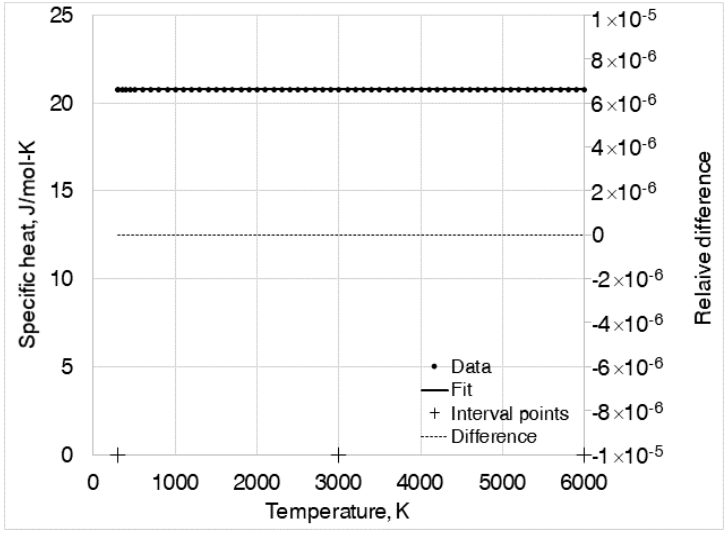


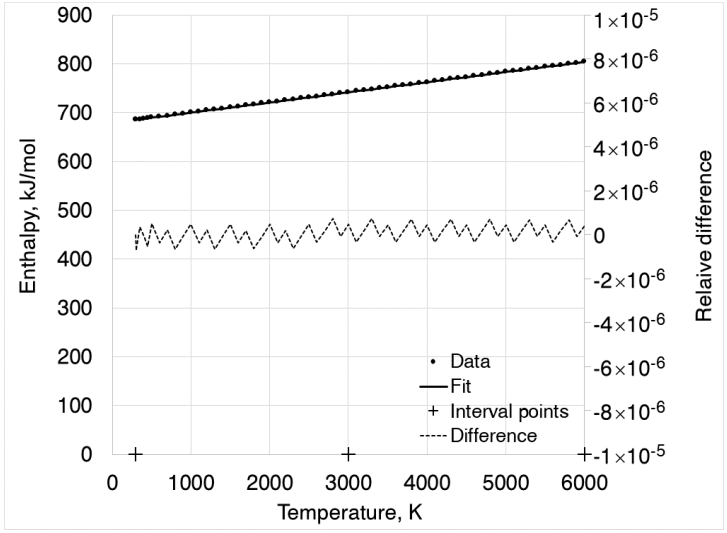


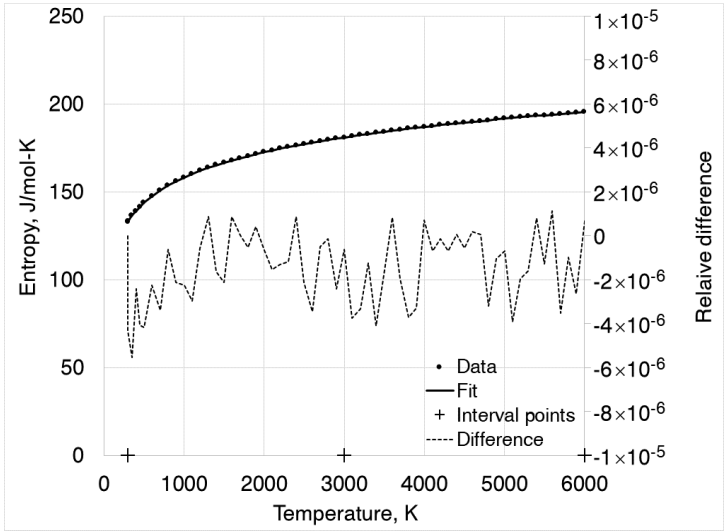


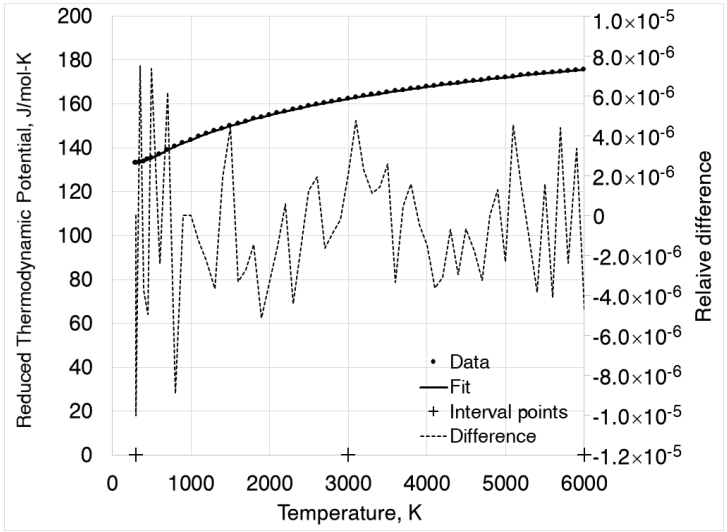
Lithium, Ion (Li^+) – Li1+(g)

MW = 6.94 g/mol – $\Delta_f H^o$(298.15 K) = 685.719 kJ/mol – S^o(298.15 K) = 133.017 J/mol-K – p^o = 0.1 MPa



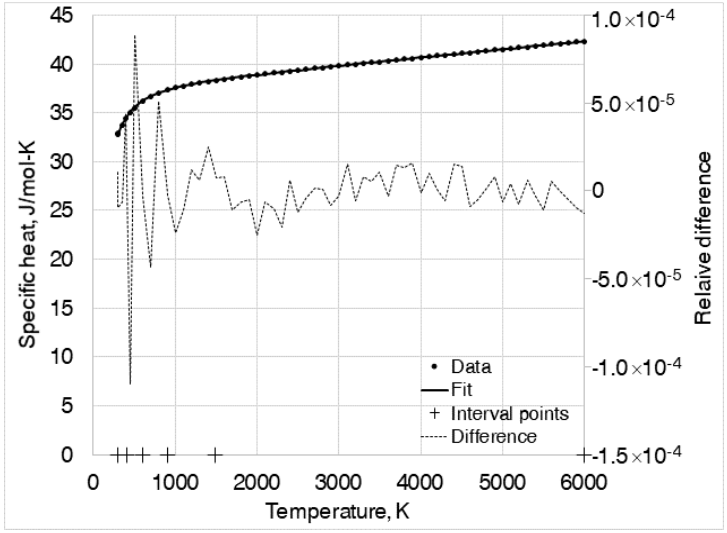


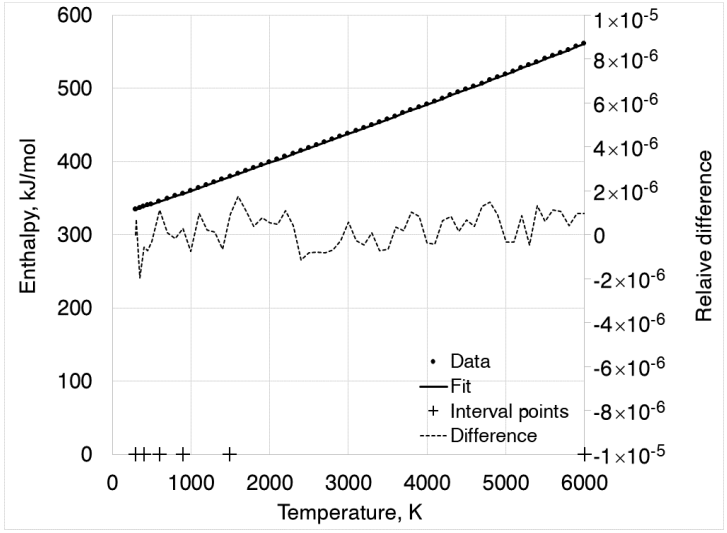


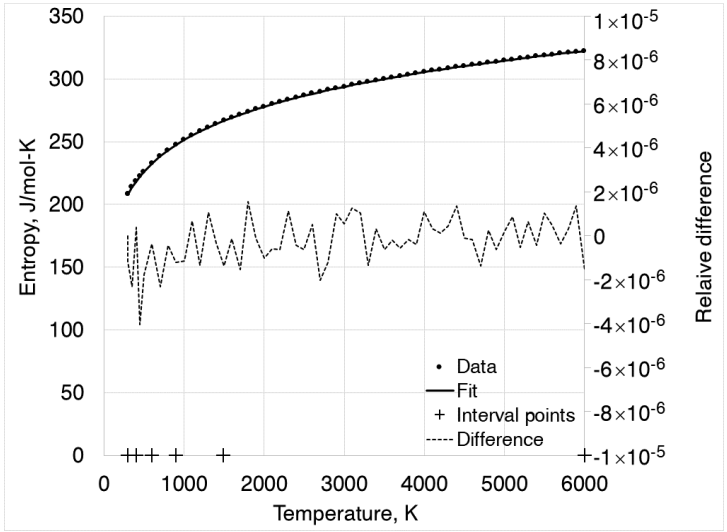


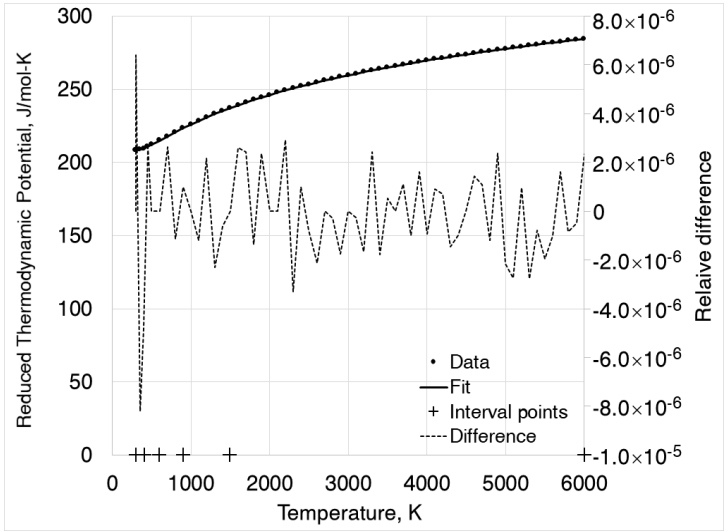
Lithium Nitride (LiN) – Li1N1(g)

MW = 20.948 g/mol – $\Delta_f H^o$(298.15 K) = 334.72 kJ/mol – S^o(298.15 K) = 208.245 J/mol-K – p^o = 0.1 MPa



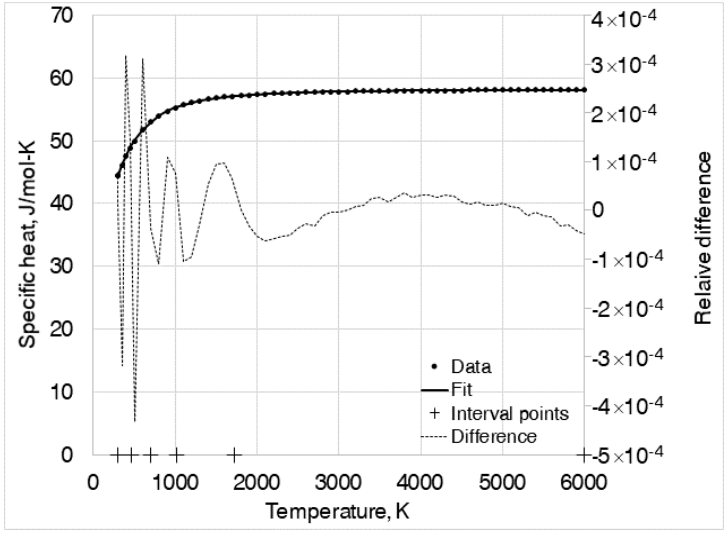


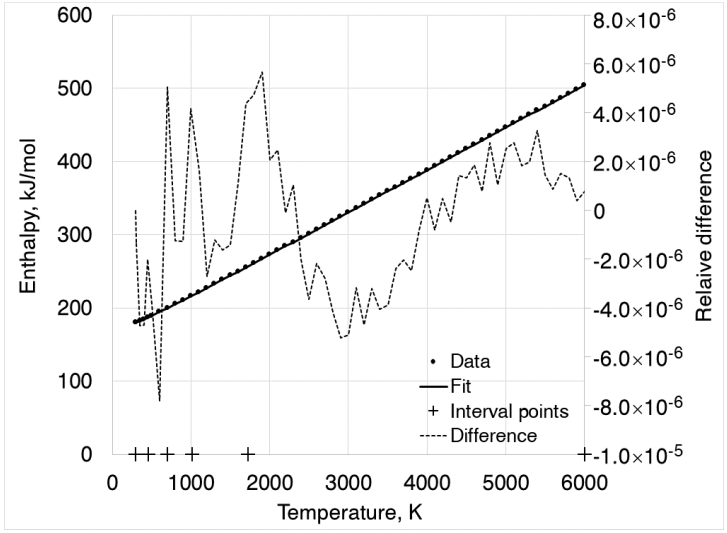


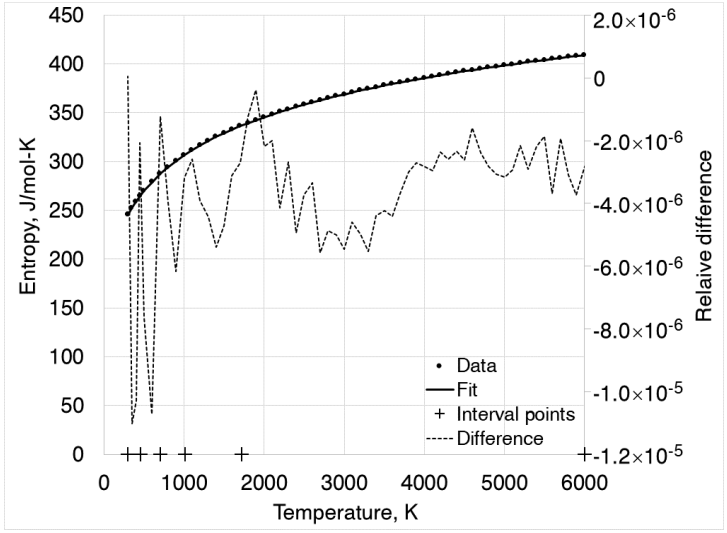


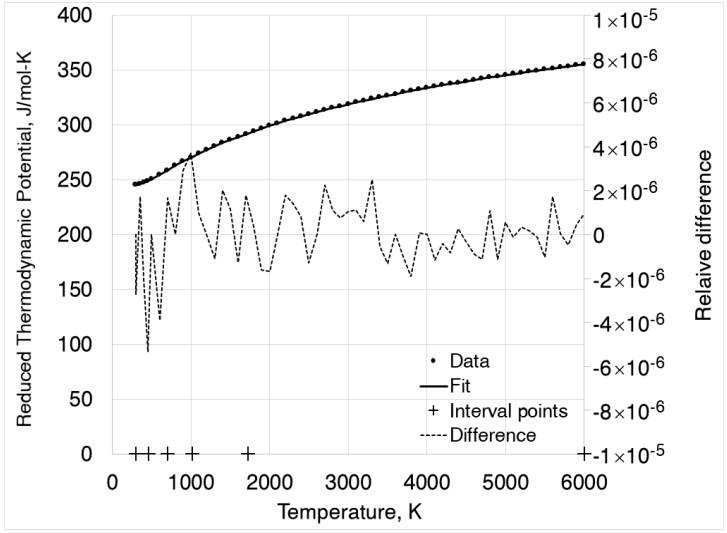
Lithium Oxynitride (LiON) – Li1N1O1(g)

MW = 36.947 g/mol – $\Delta_f H^o$(298.15 K) = 179.912 kJ/mol – S^o(298.15 K) = 245.332 J/mol-K – p^o = 0.1 MPa



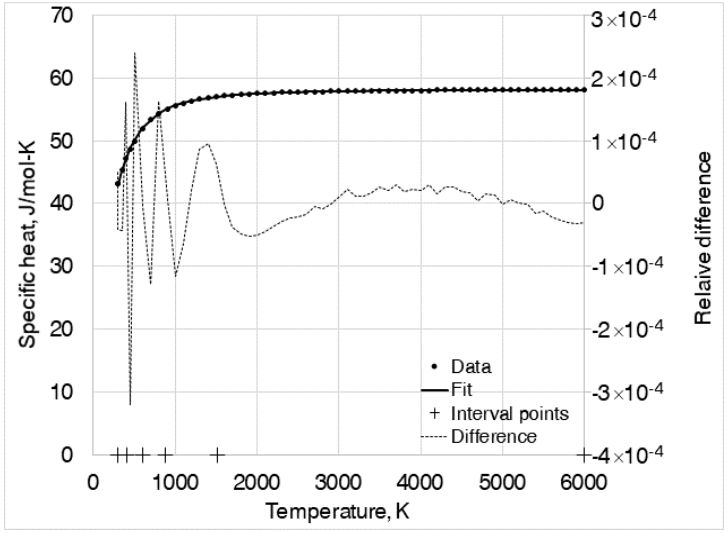


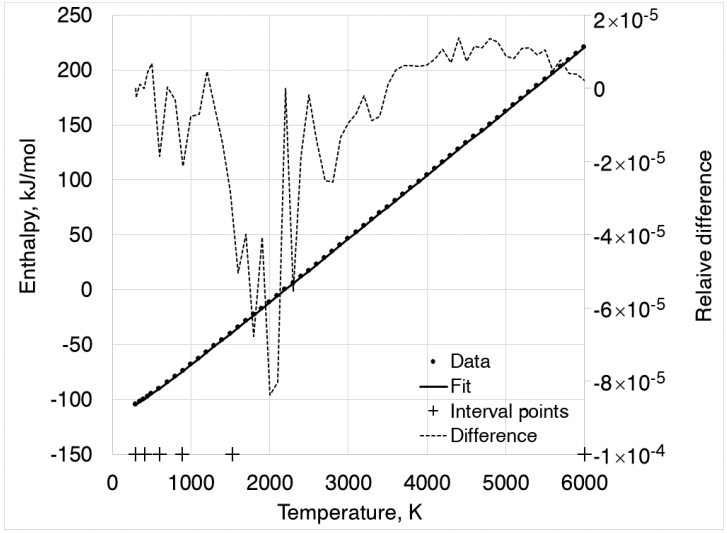


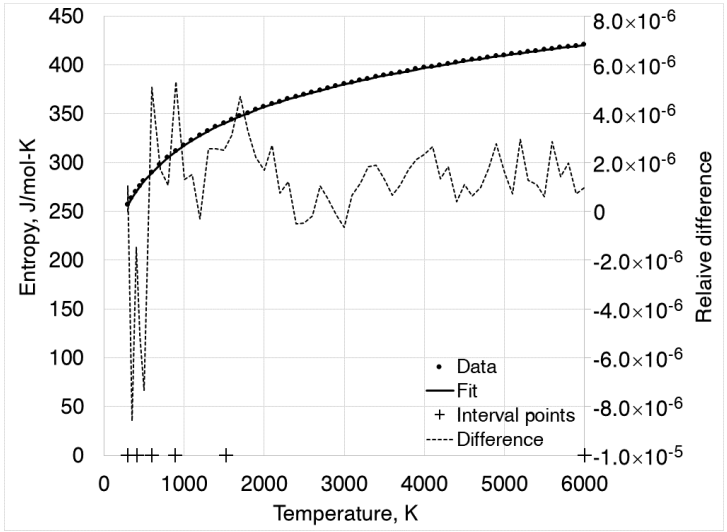


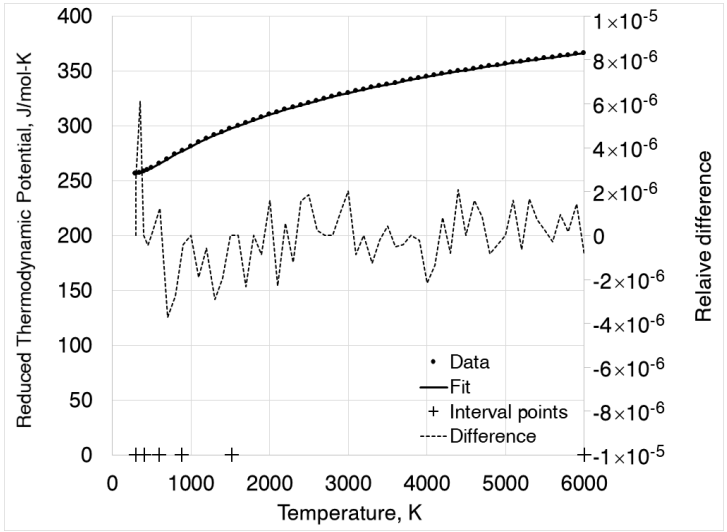
Lithium Sodium Oxide (LiONa) – Li1Na1O1(g)

MW = 45.93 g/mol – $\Delta_f H^o$(298.15 K) = -104.6 kJ/mol – S^o(298.15 K) = 256.459 J/mol-K – p^o = 0.1 MPa



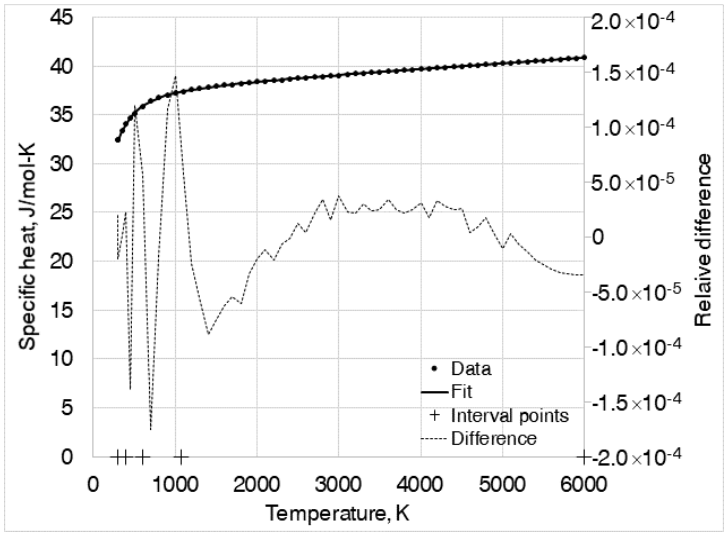


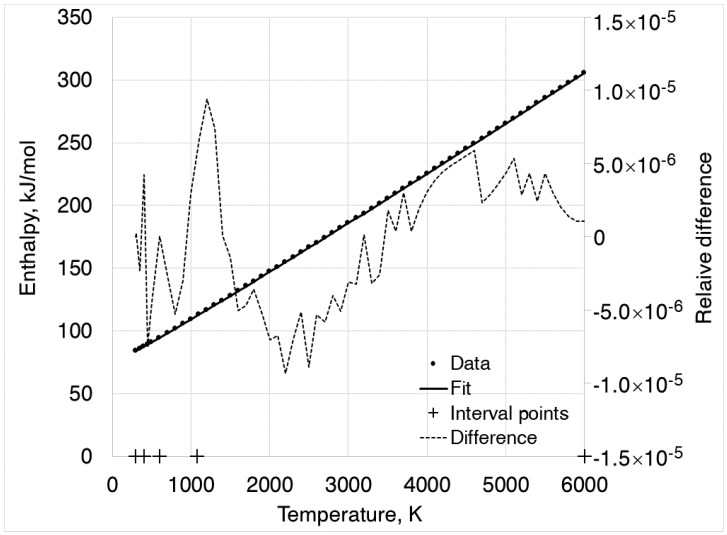


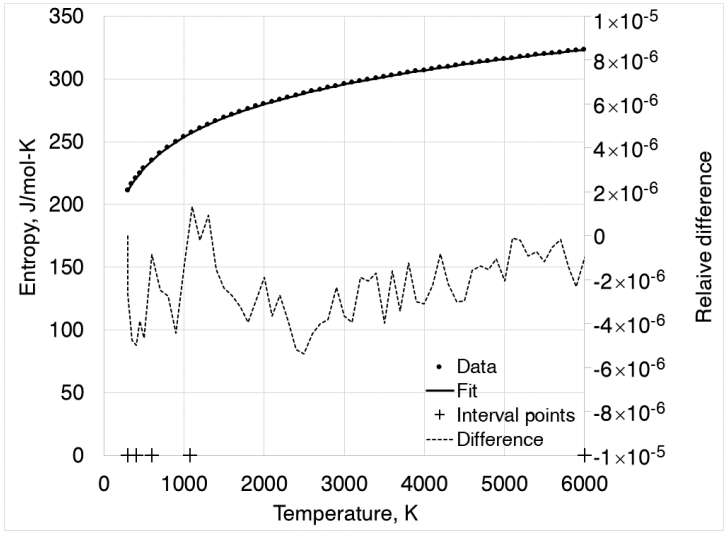


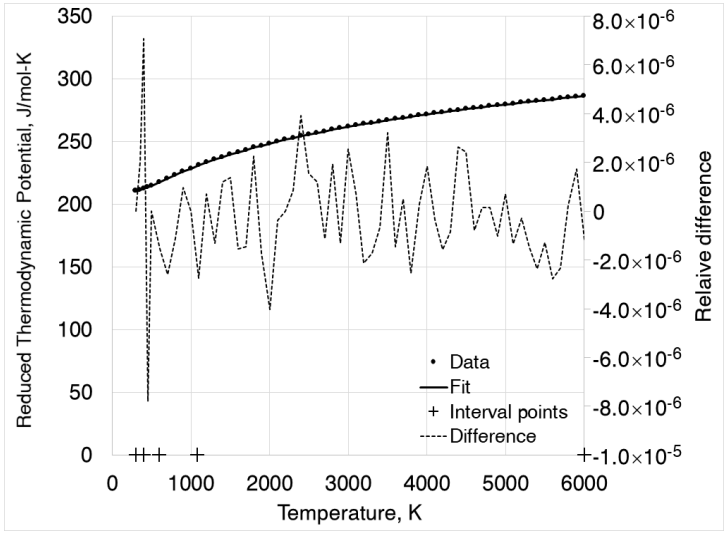
Lithium Oxide (LiO) – Li1O1(g)

MW = 22.94 g/mol – $\Delta_f H^o$(298.15 K) = 84.098 kJ/mol – S^o(298.15 K) = 210.963 J/mol-K – p^o = 0.1 MPa



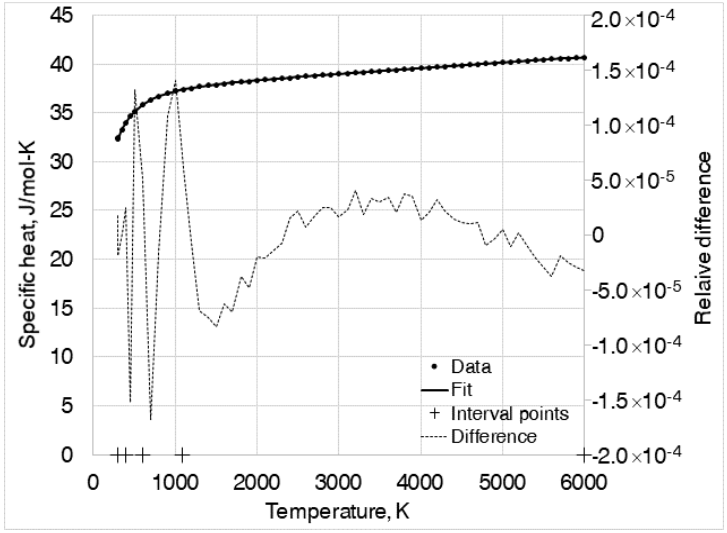


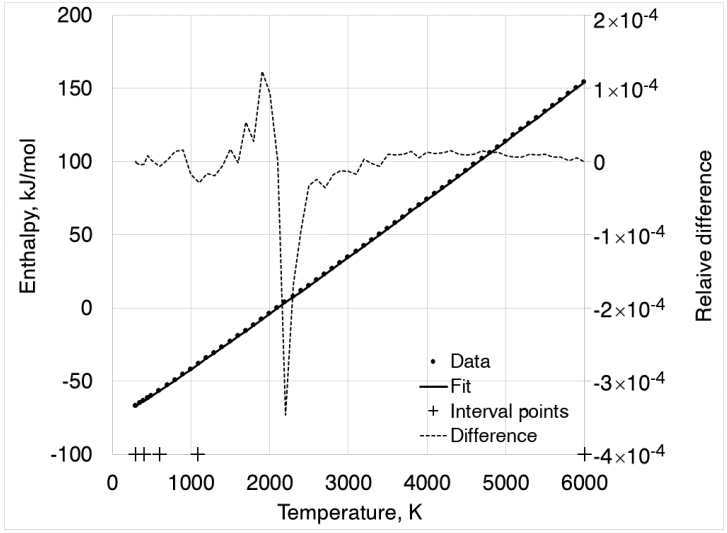


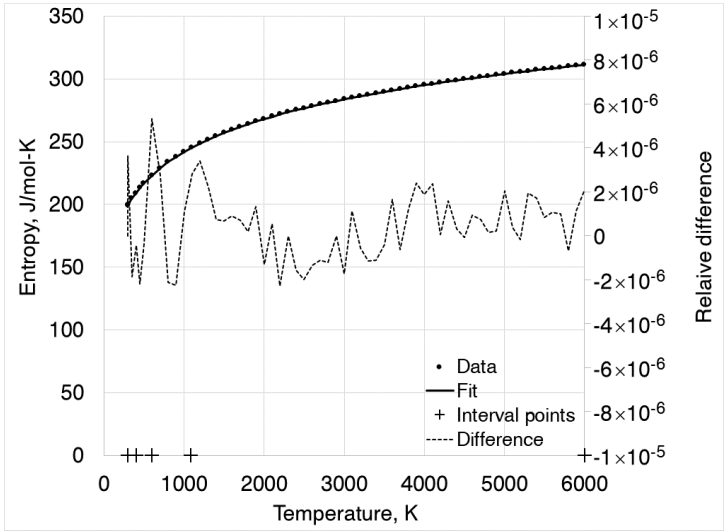


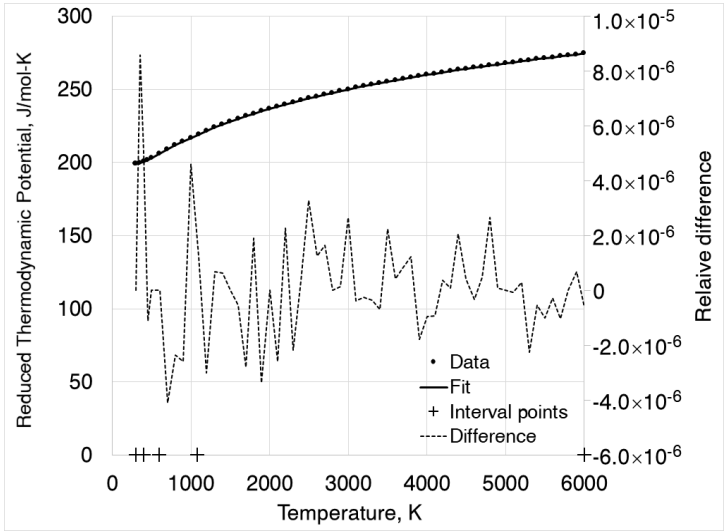
Lithium Oxide, Ion (LiO^-) – Li1O1-(g)

MW = 22.941 g/mol – $\Delta_f H^o$(298.15 K) = -66.944 kJ/mol – S^o(298.15 K) = 199.241 J/mol-K – p^o = 0.1 MPa



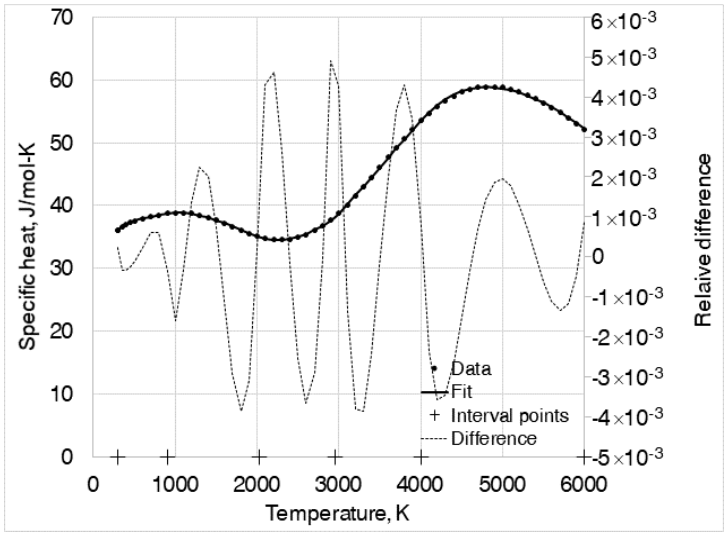


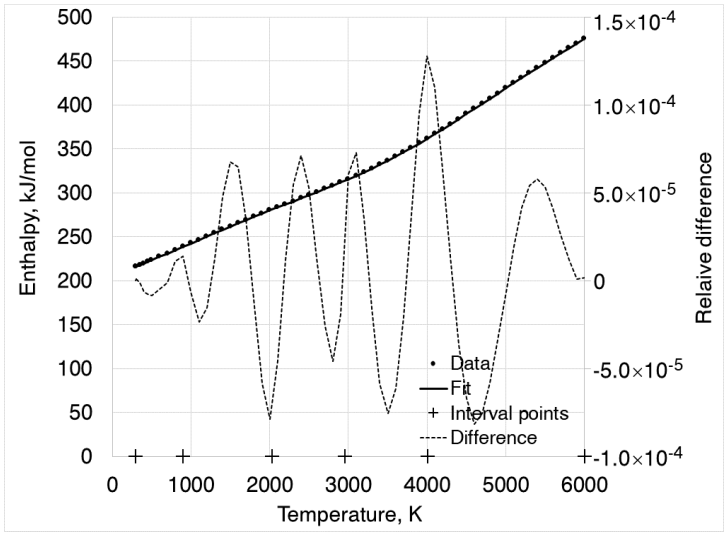


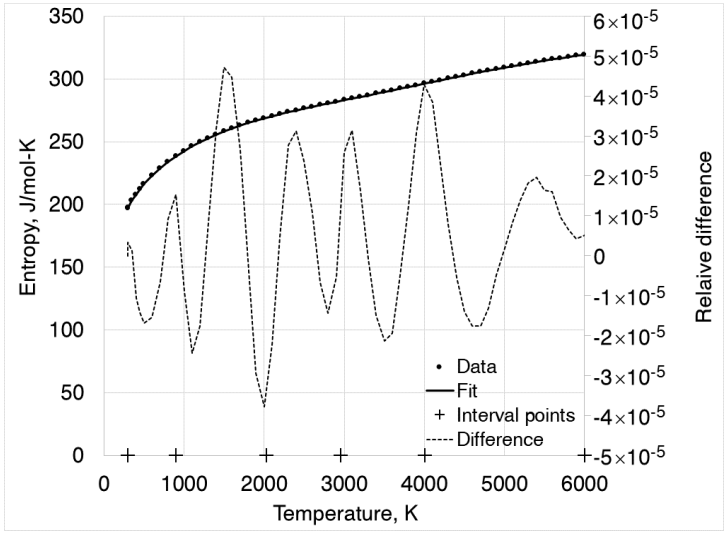


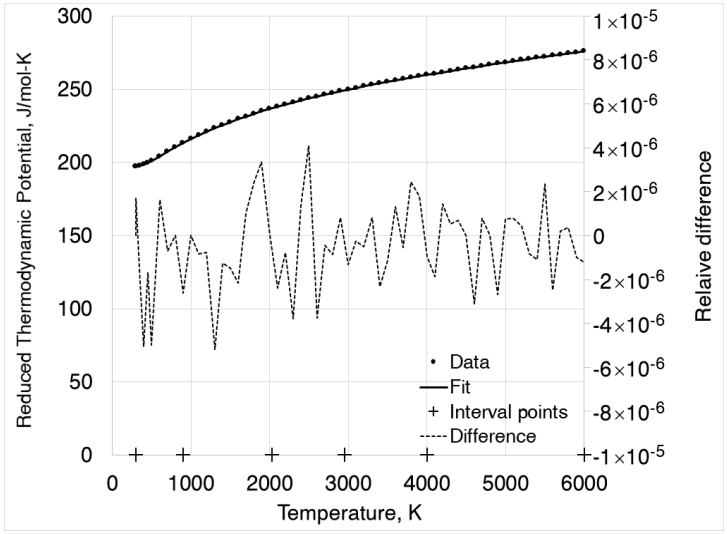
Lithium (Li_2) – Li2(g)

MW = 13.882 g/mol – $\Delta_f H^o$(298.15 K) = 215.9 kJ/mol – S^o(298.15 K) = 196.998 J/mol-K – p^o = 0.1 MPa



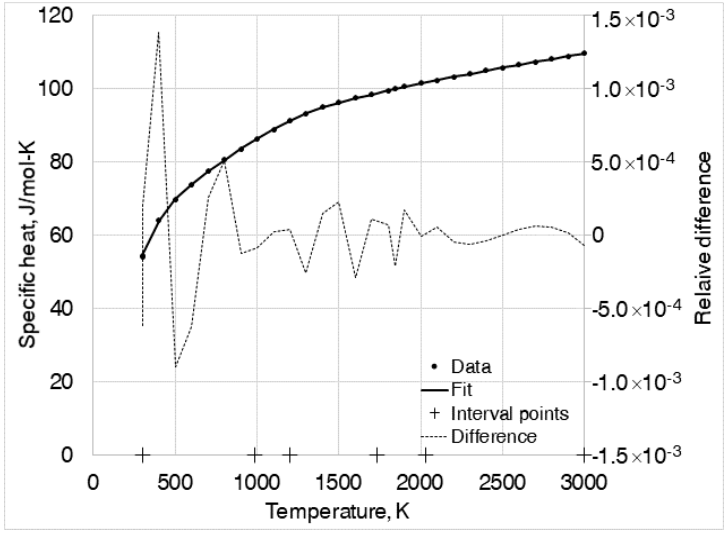


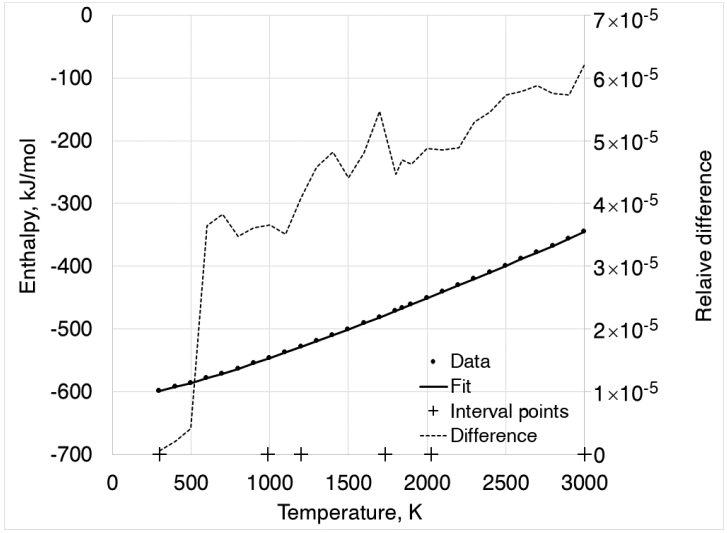


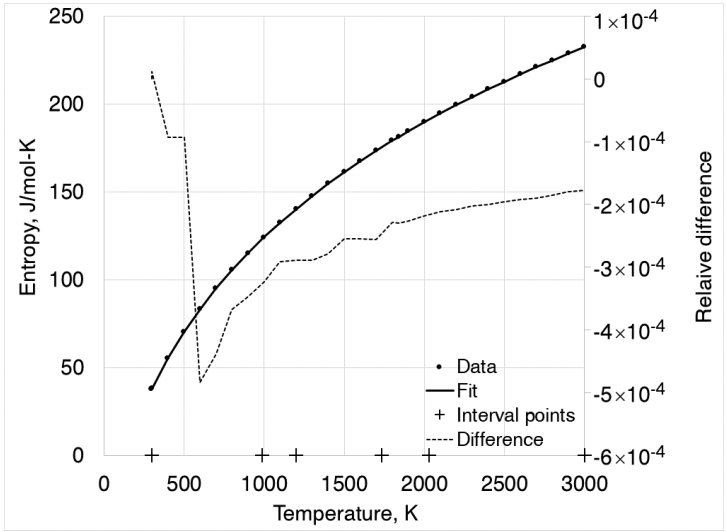


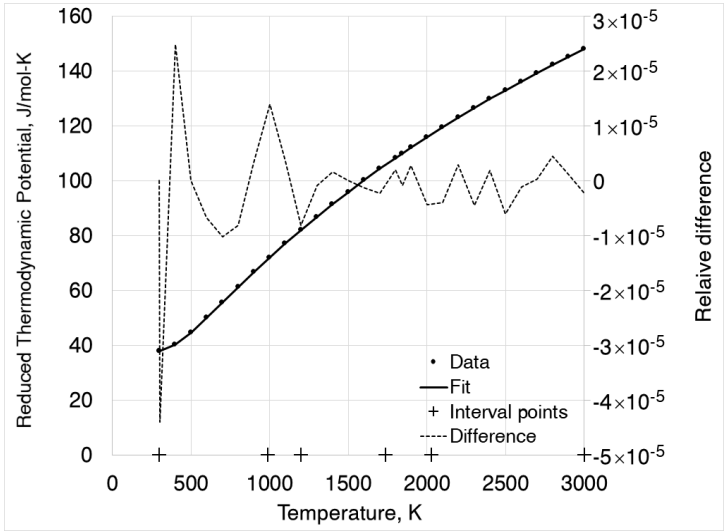
Lithium Oxide (Li_2O) – Li2O1(cr)

MW = 29.881 g/mol – $\Delta_f H^o$(298.15 K) = -598.73 kJ/mol – S^o(298.15 K) = 37.89 J/mol-K – p^o = 0.1 MPa



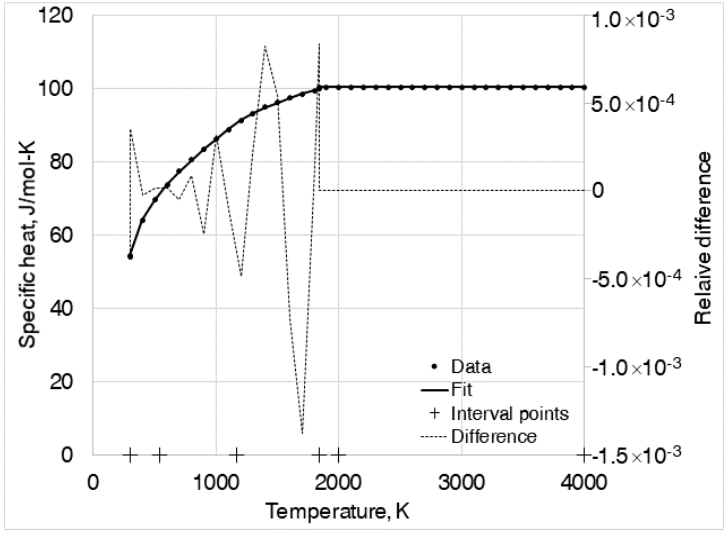


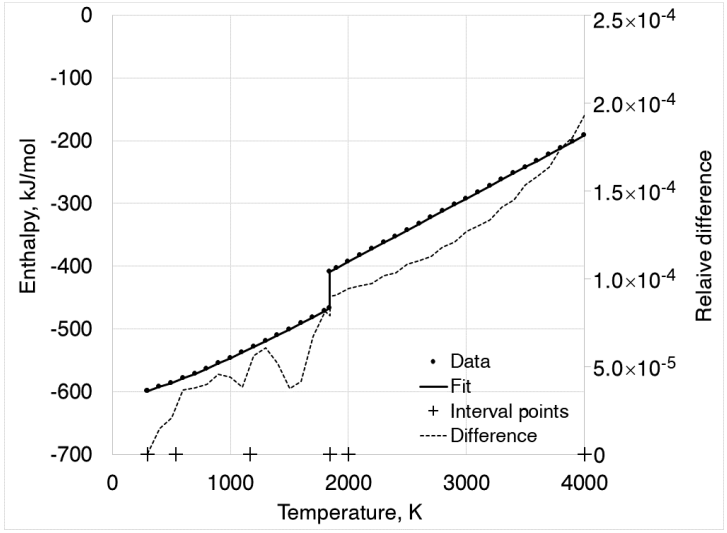


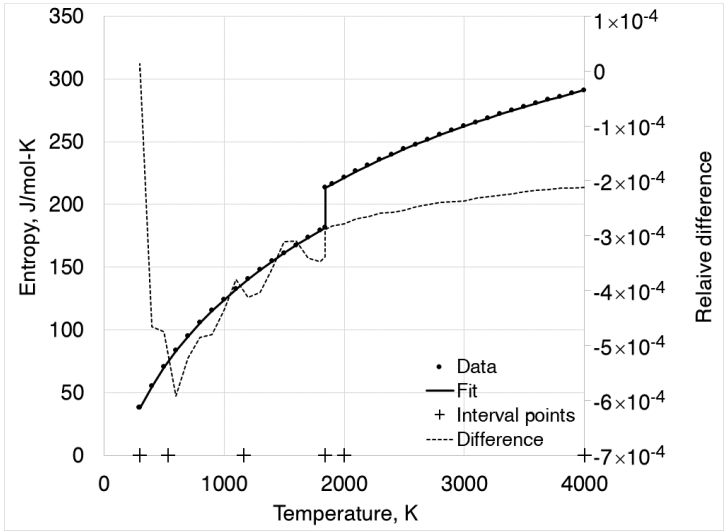


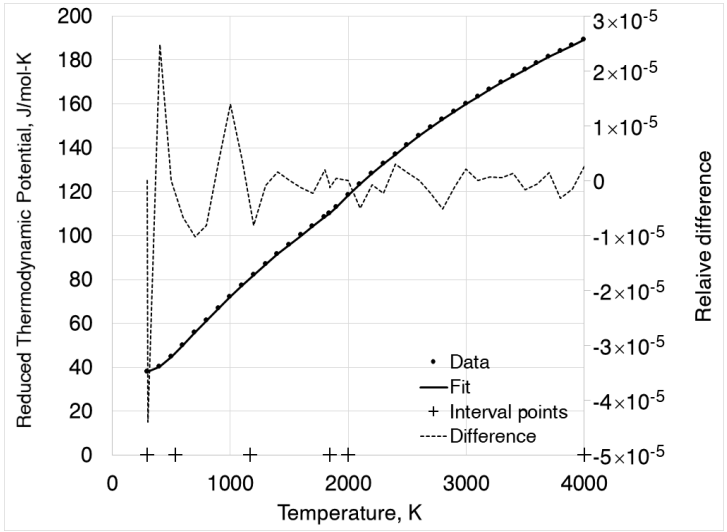
Lithium Oxide (Li_2O) – Li2O1(cr,l)

MW = 29.881 g/mol – $\Delta_f H^o$(298.15 K) = -598.73 kJ/mol – S^o(298.15 K) = 37.89 J/mol-K – p^o = 0.1 MPa



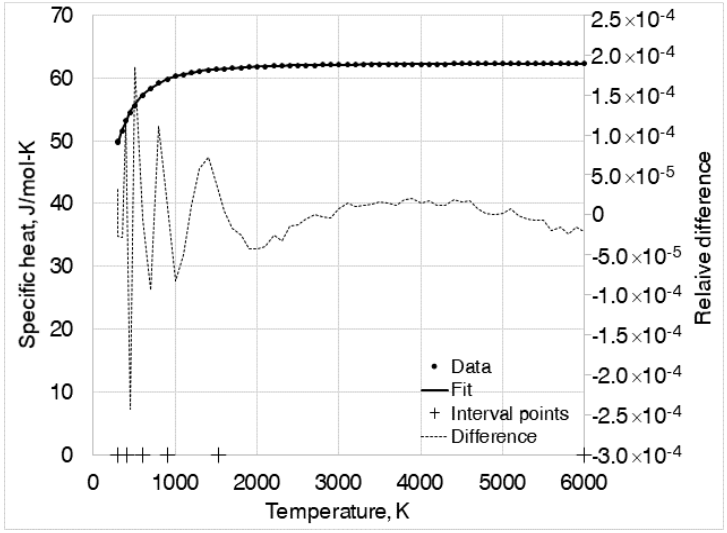


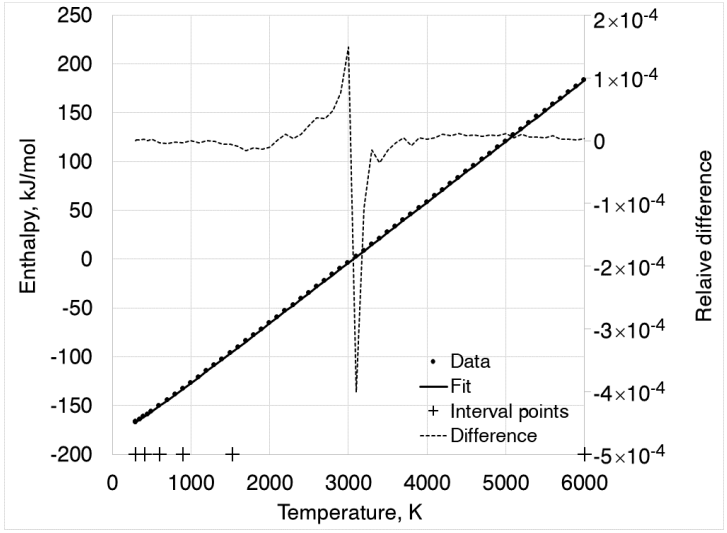


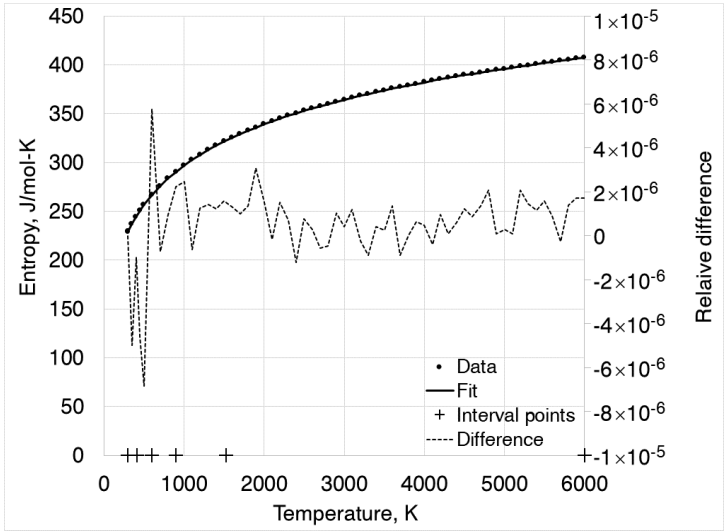


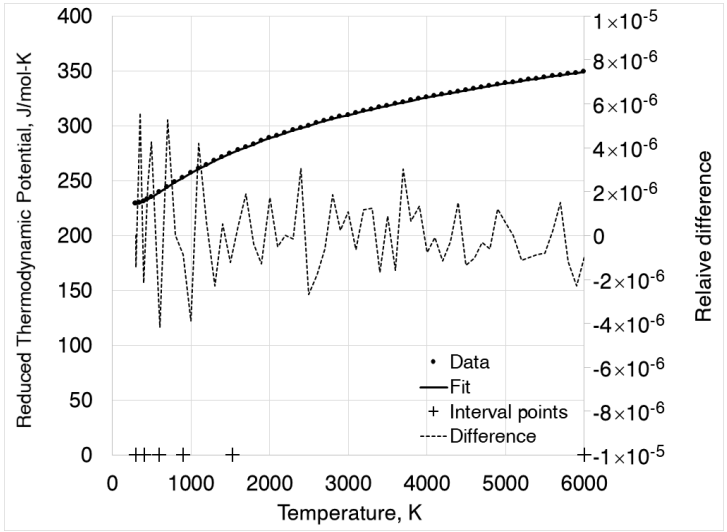
Lithium Oxide (Li_2O) – Li2O1(g)

MW = 29.881 g/mol – $\Delta_f H^o$(298.15 K) = -166.942 kJ/mol – S^o(298.15 K) = 229.108 J/mol-K – p^o = 0.1 MPa



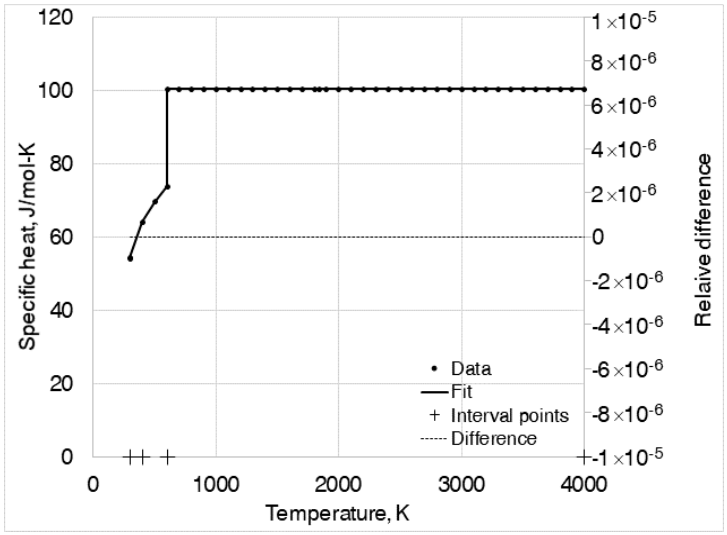


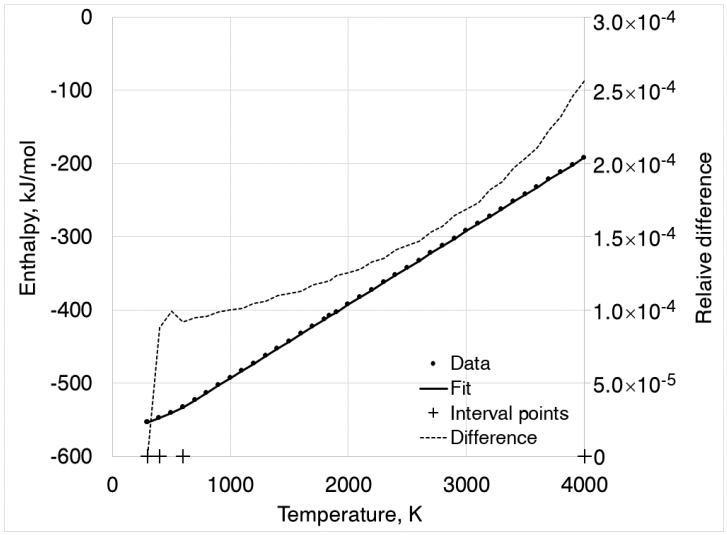


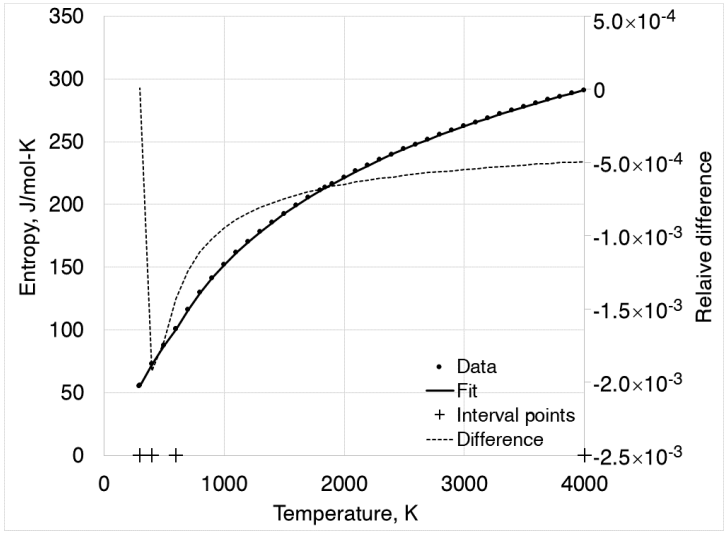


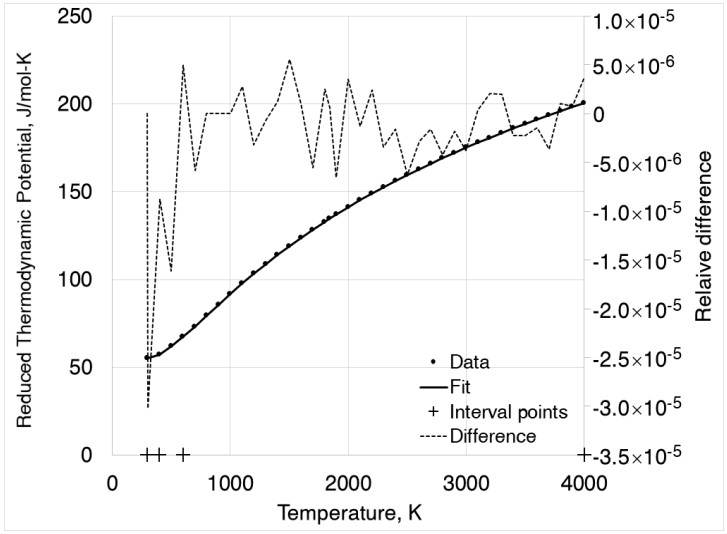
Lithium Oxide (Li_2O) – Li2O1(l)

MW = 29.881 g/mol – $\Delta_f H^o$(298.15 K) = -553.247 kJ/mol – S^o(298.15 K) = 55.083 J/mol-K – p^o = 0.1 MPa



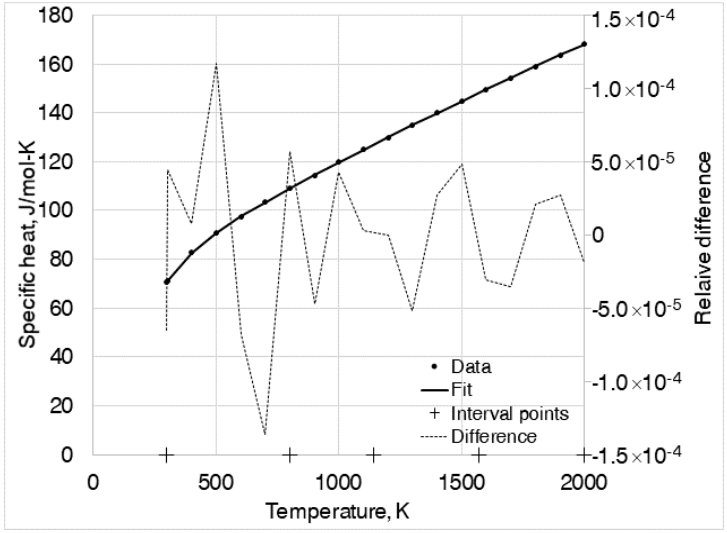


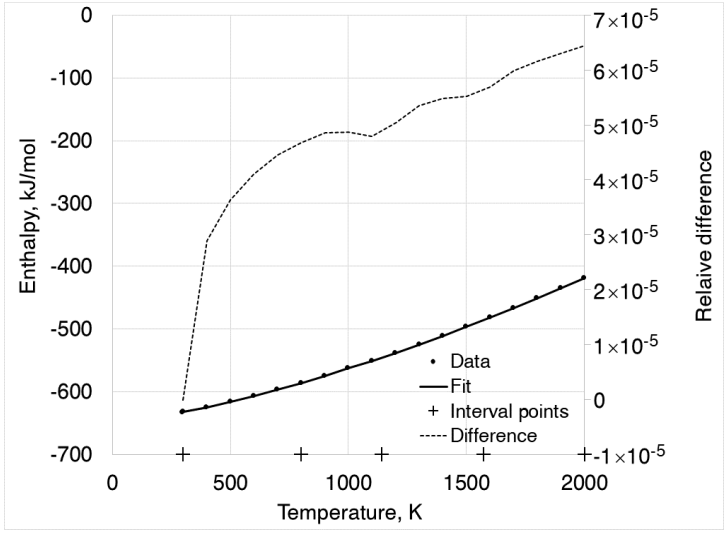


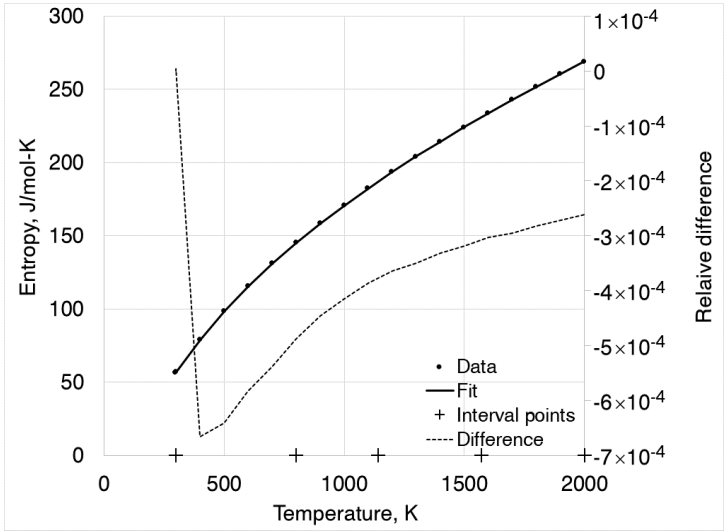


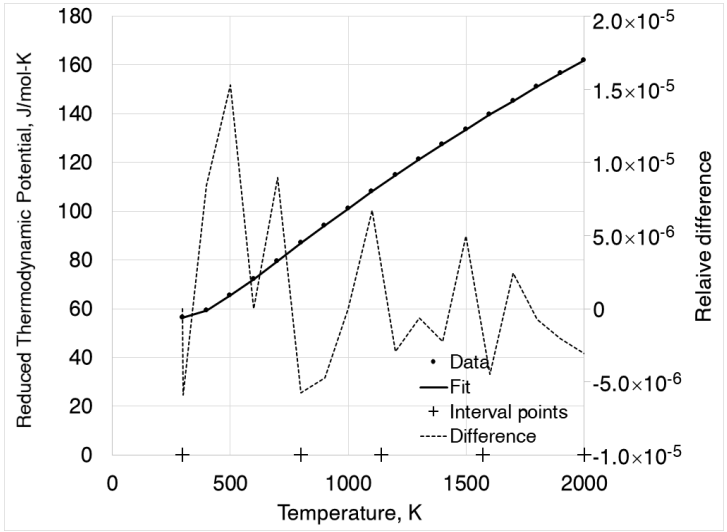
Lithium Peroxide (Li_2O_2) – Li2O2(cr)

MW = 45.881 g/mol – $\Delta_f H^o$(298.15 K) = -632.621 kJ/mol – S^o(298.15 K) = 56.484 J/mol-K – p^o = 0.1 MPa



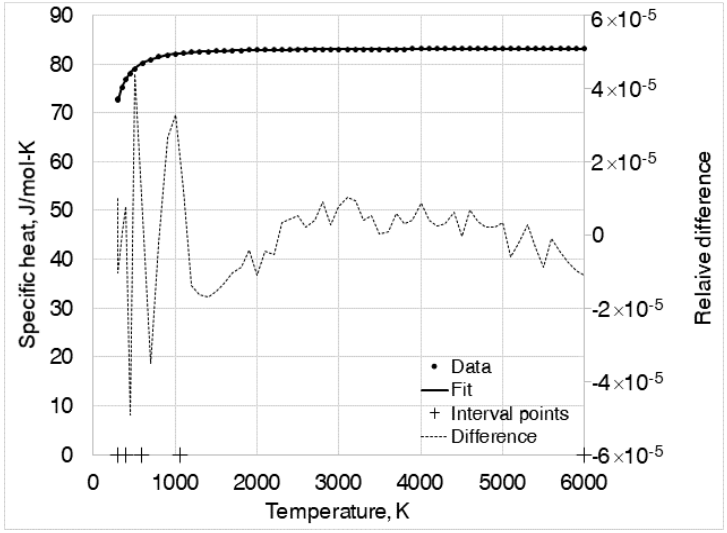


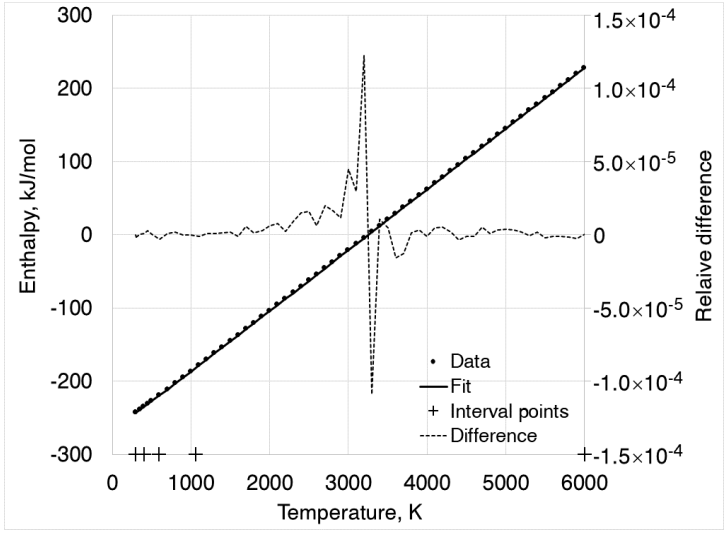


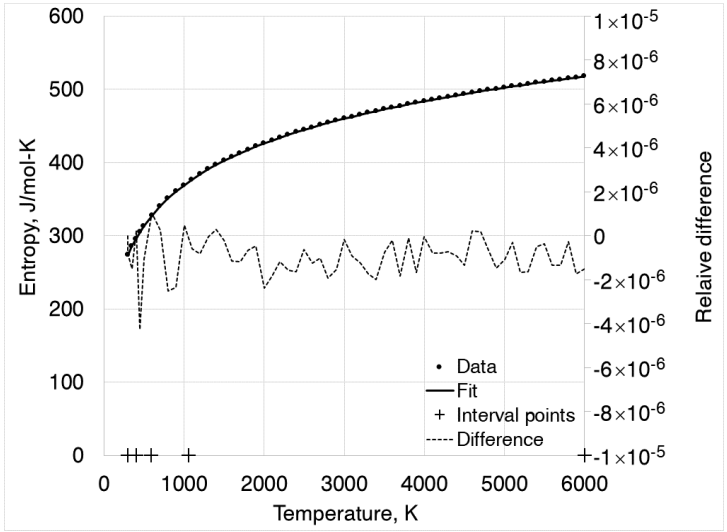


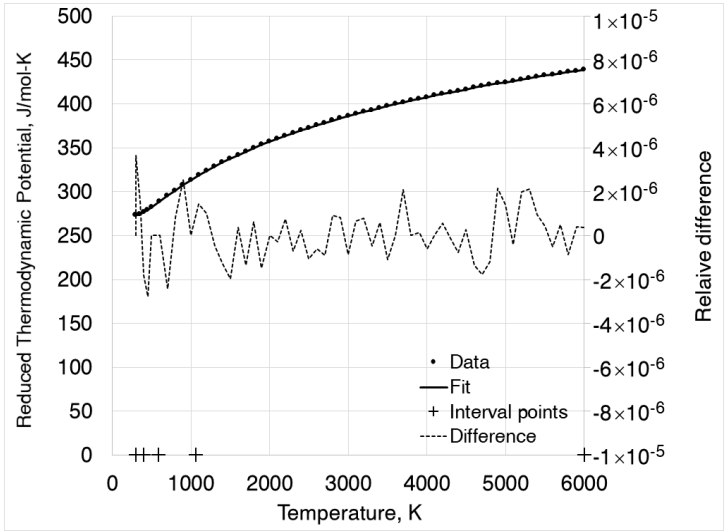
Lithium Oxide ((LiO)$_2$) – Li2O2(g)

MW = 45.881 g/mol – $\Delta_f H^o$(298.15 K) = -242.672 kJ/mol – S^o(298.15 K) = 273.562 J/mol-K – p^o = 0.1 MPa



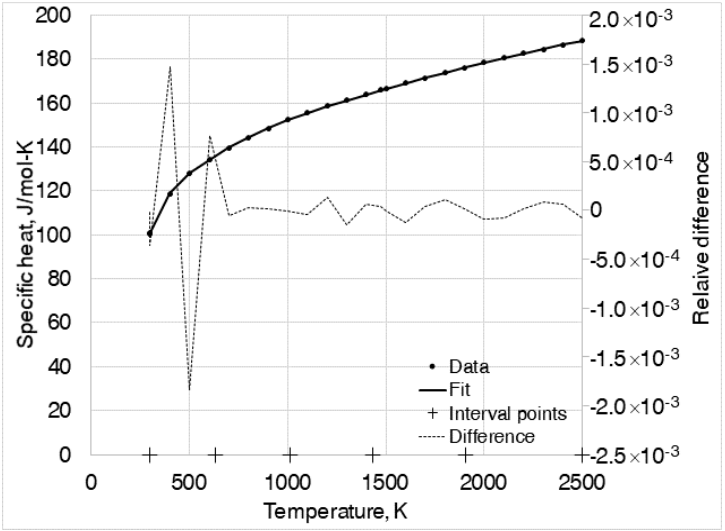


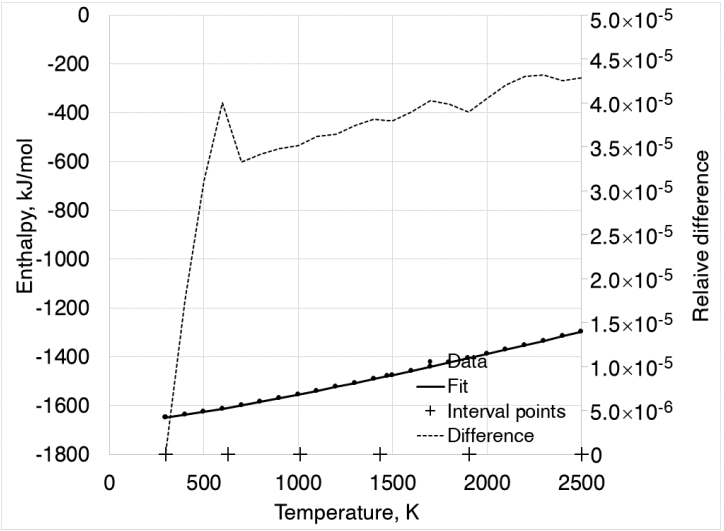


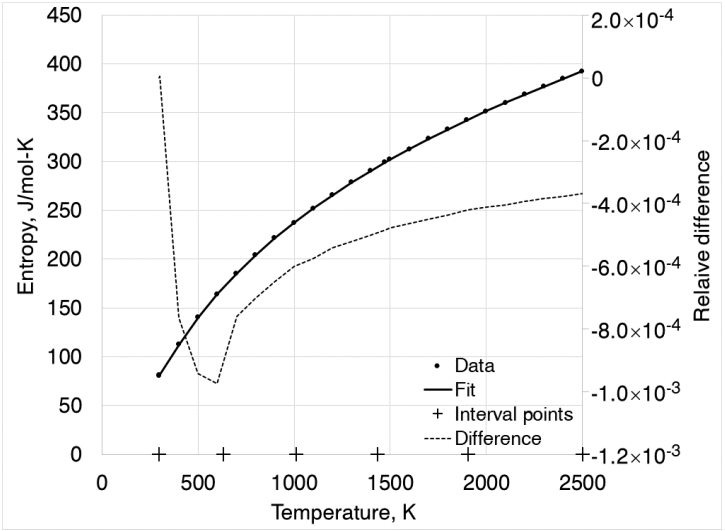


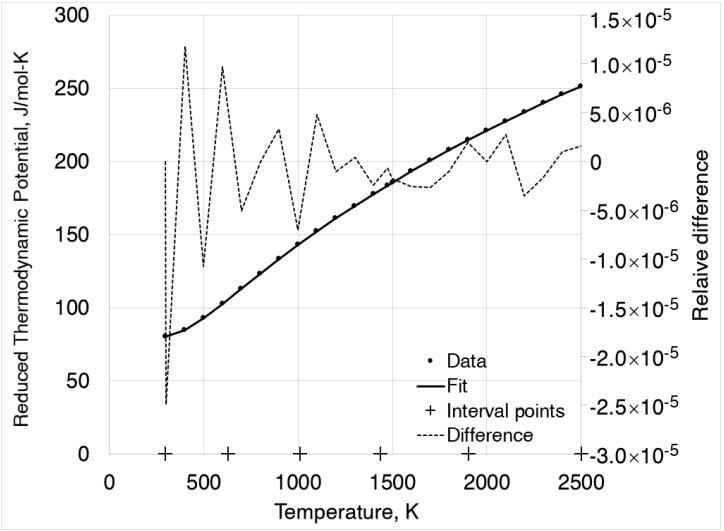
Lithium Silicate (Li_2SiO_3) – Li2O3Si1(cr)

MW = 89.966 g/mol – $\Delta_f H^o$(298.15 K) = -1649.5 kJ/mol – S^o(298.15 K) = 80.295 J/mol-K – p^o = 0.1 MPa



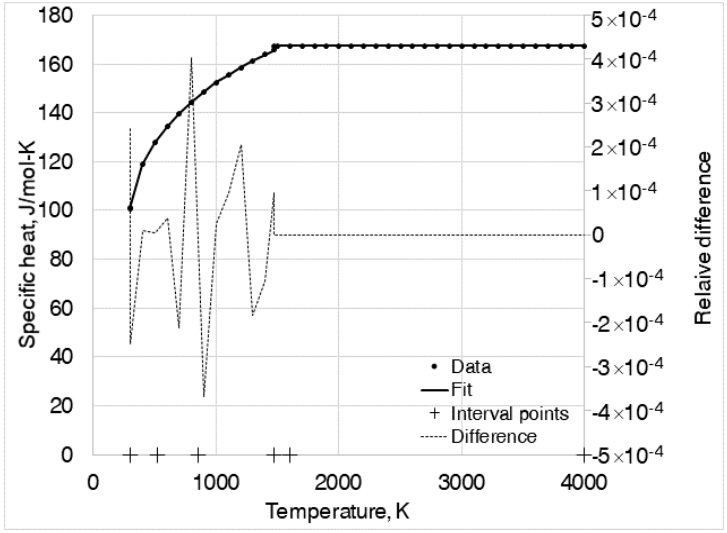


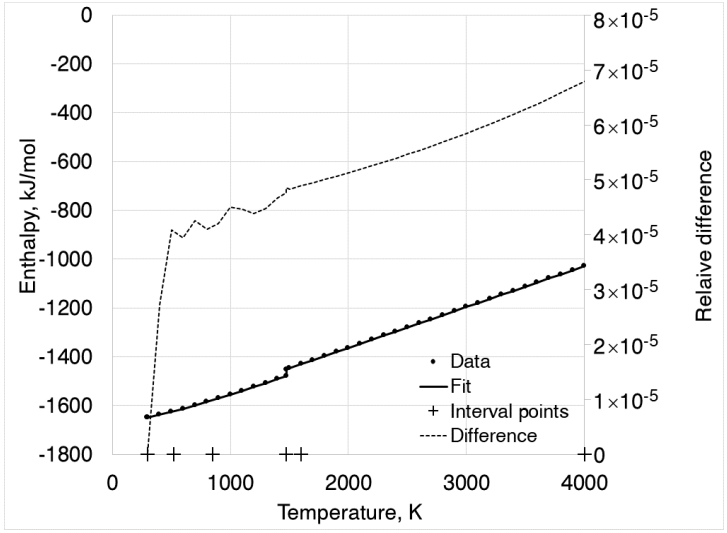


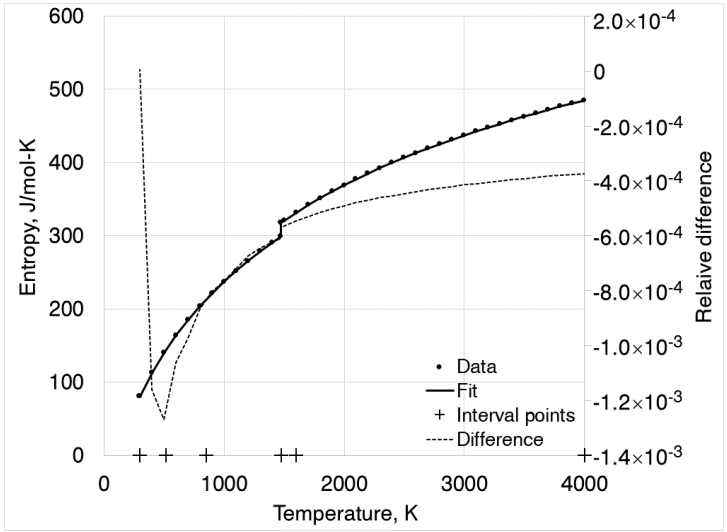


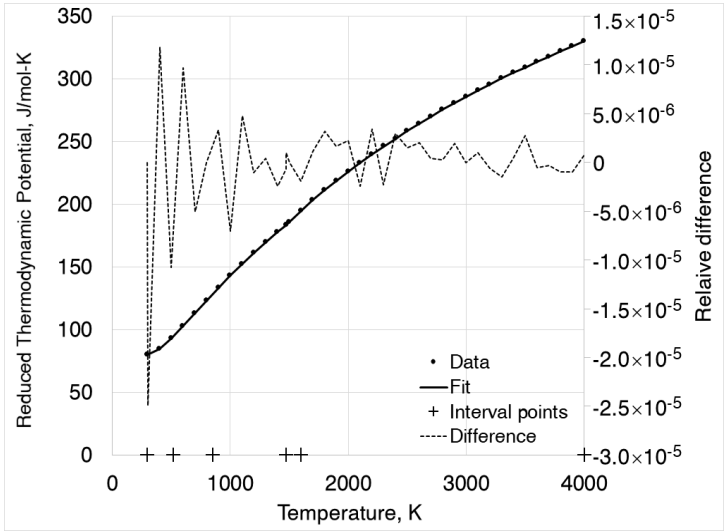
Lithium Silicate (Li_2SiO_3) – Li2O3Si1(cr,l)

MW = 89.966 g/mol – $\Delta_f H^o$(298.15 K) = -1649.5 kJ/mol – S^o(298.15 K) = 80.295 J/mol-K – p^o = 0.1 MPa



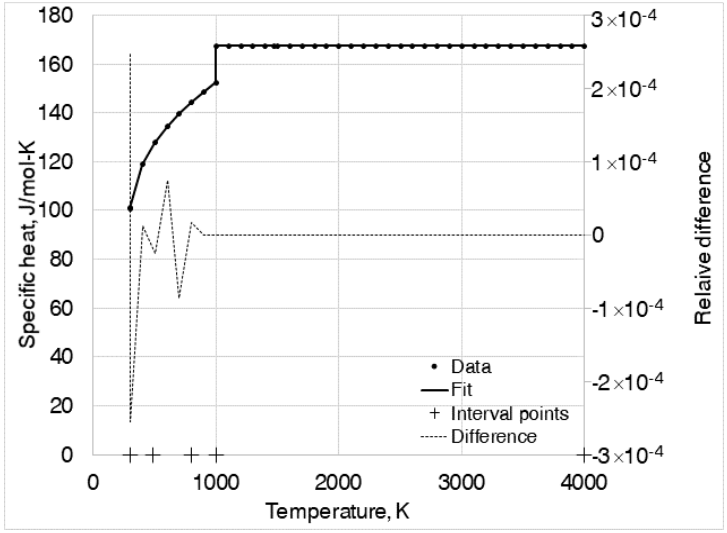


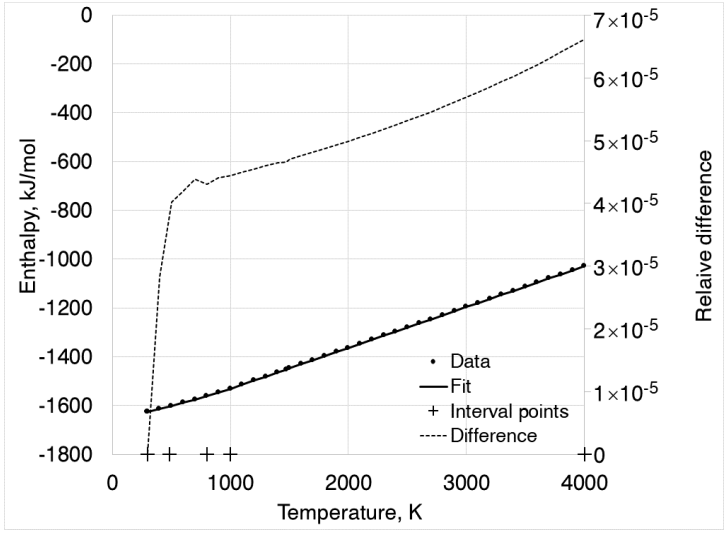


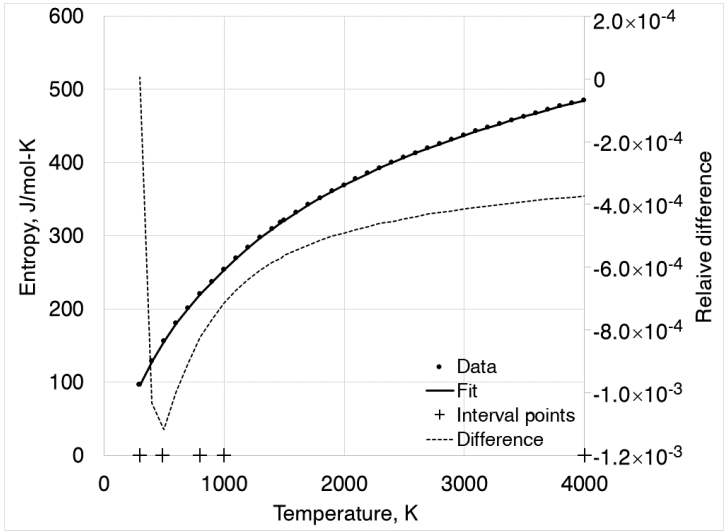


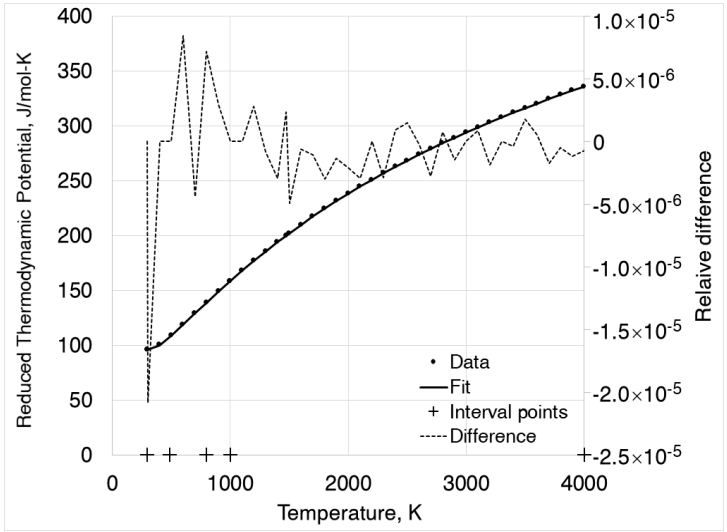
Lithium Silicate (Li_2SiO_3) – Li2O3Si1(l)

MW = 89.966 g/mol – $\Delta_f H^o$(298.15 K) = -1625.231 kJ/mol – S^o(298.15 K) = 96.064 J/mol-K – p^o = 0.1 MPa



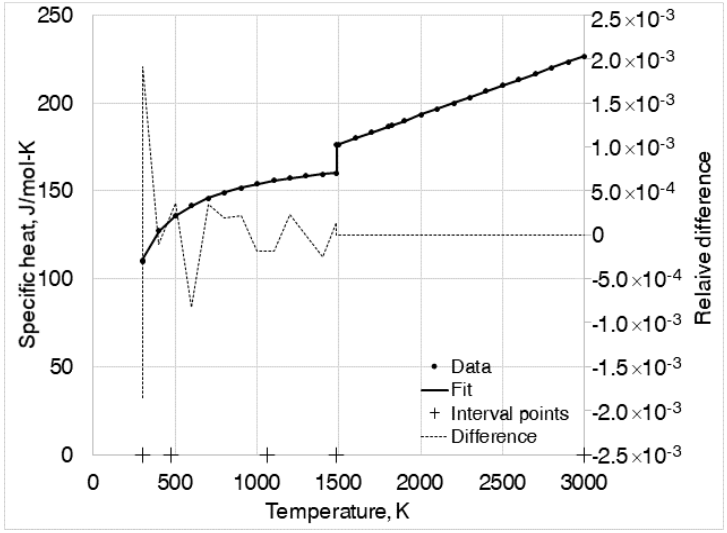


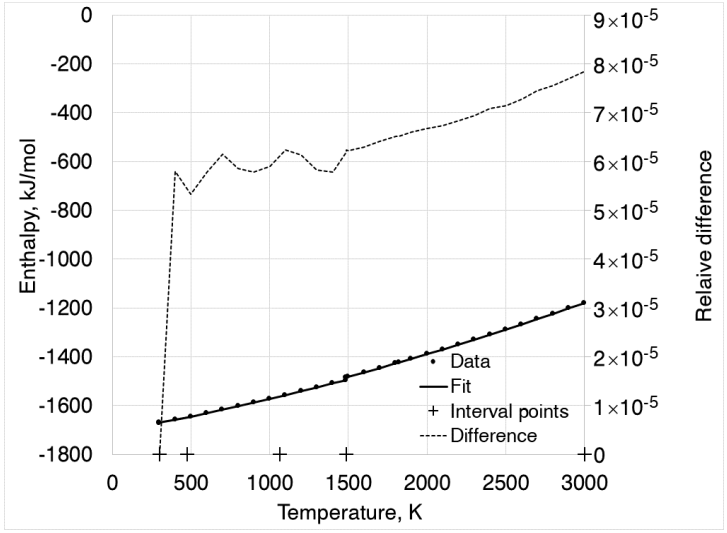


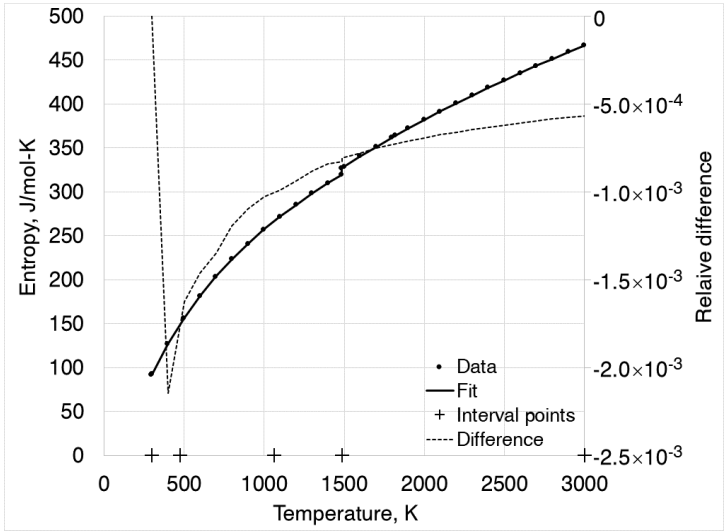


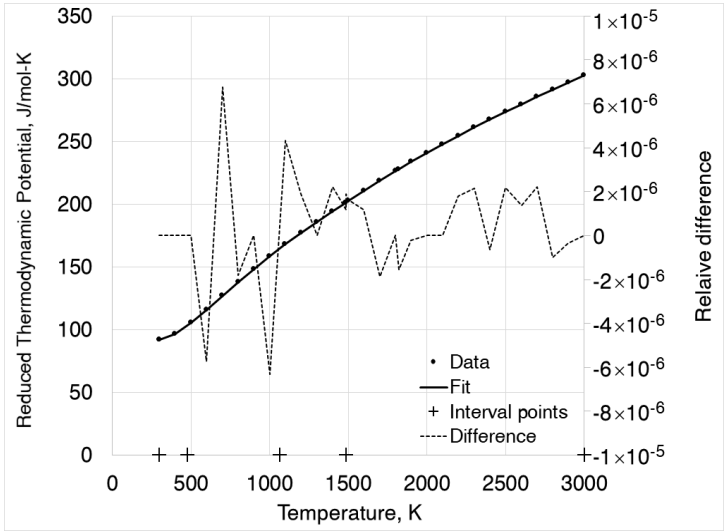
Lithium Titanium Oxide (Li_2TiO_3) – Li2O3Ti1(cr)

MW = 109.76 g/mol – $\Delta_f H^o$(298.15 K) = -1670.671 kJ/mol – S^o(298.15 K) = 91.755 J/mol-K – p^o = 0.1 MPa



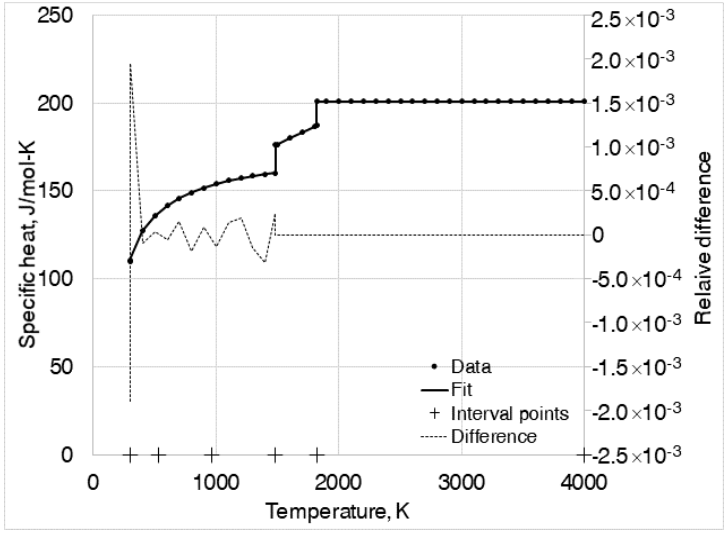


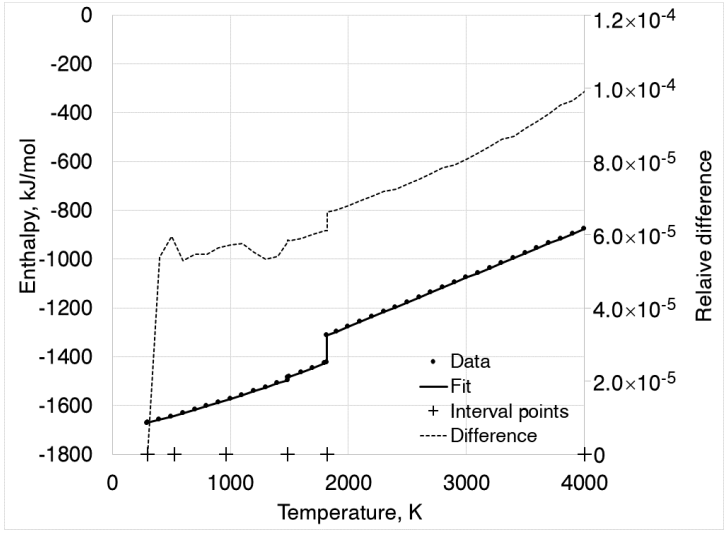


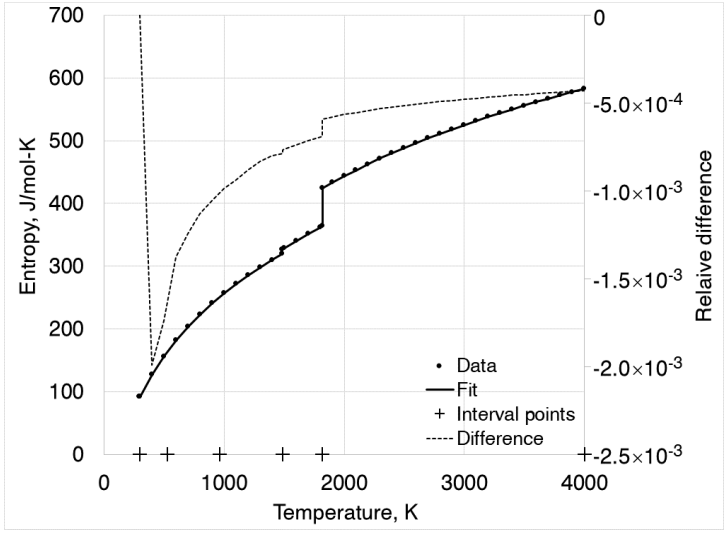


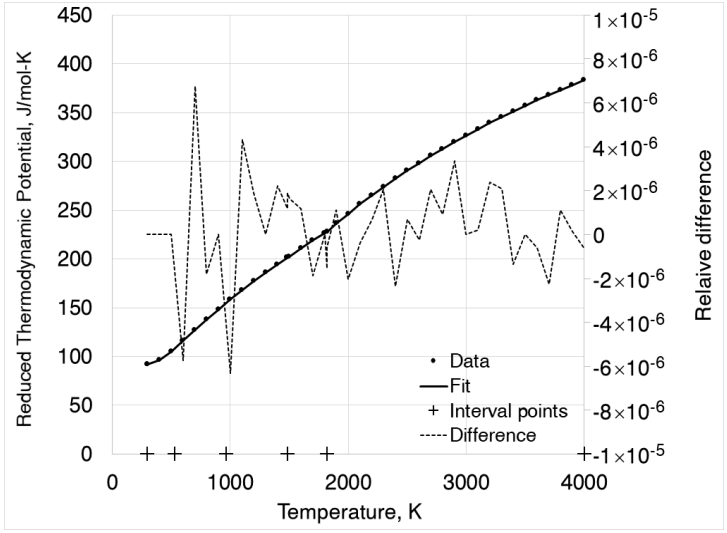
Lithium Titanium Oxide (Li_2TiO_3) – Li2O3Ti1(cr,l)

MW = 109.76 g/mol – $\Delta_f H^o$(298.15 K) = -1670.671 kJ/mol – S^o(298.15 K) = 91.755 J/mol-K – p^o = 0.1 MPa



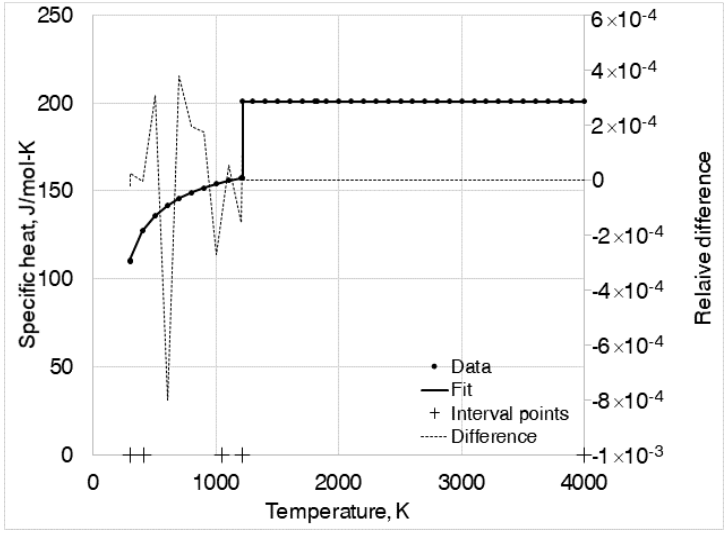


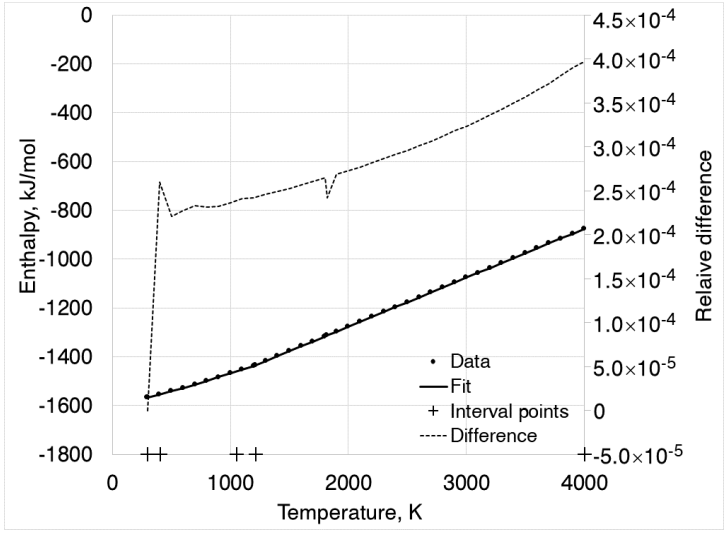


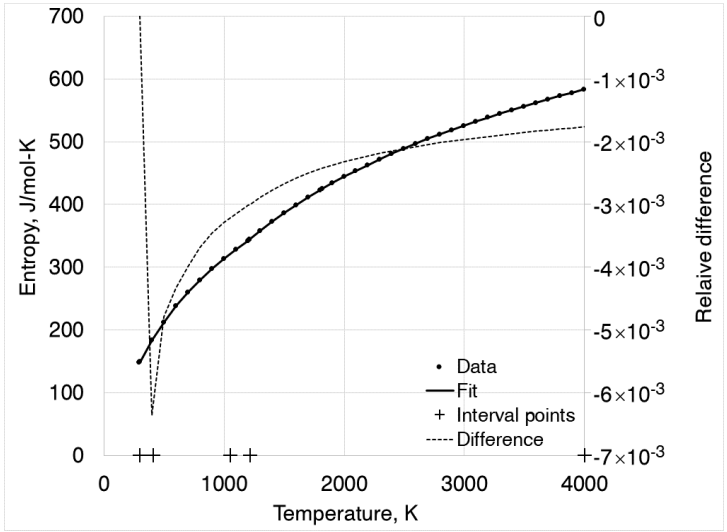


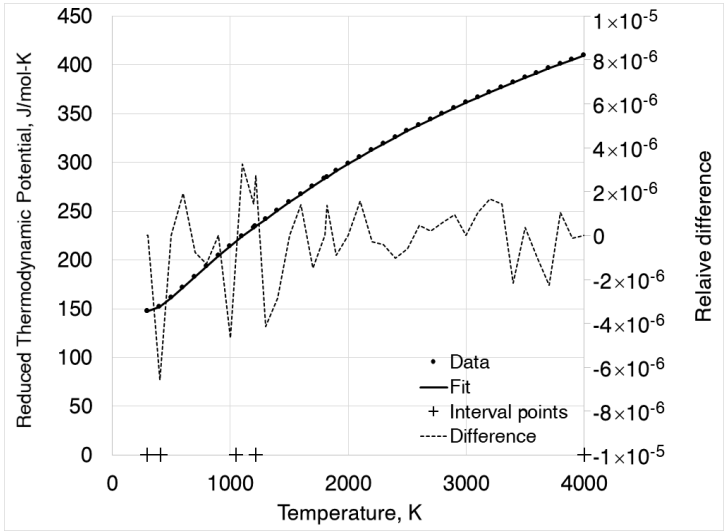
Lithium Titanium Oxide (Li_2TiO_3) – Li2O3Ti1(l)

MW = 109.76 g/mol – $\Delta_f H^o$(298.15 K) = -1566.76 kJ/mol – S^o(298.15 K) = 147.656 J/mol-K – p^o = 0.1 MPa



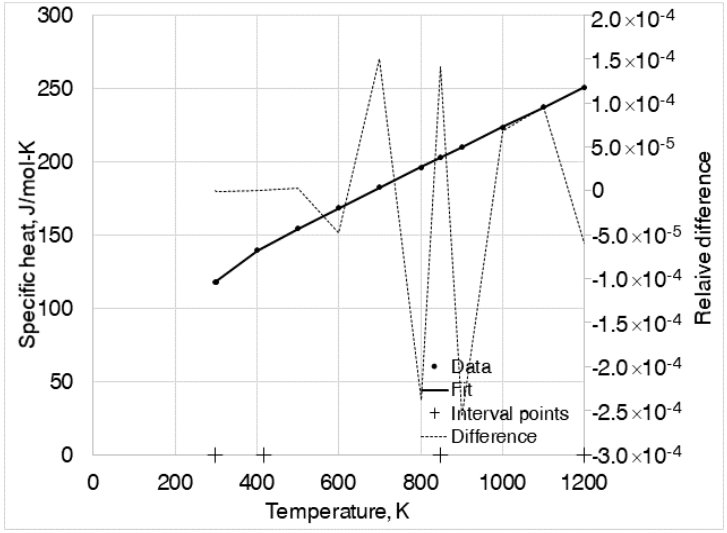


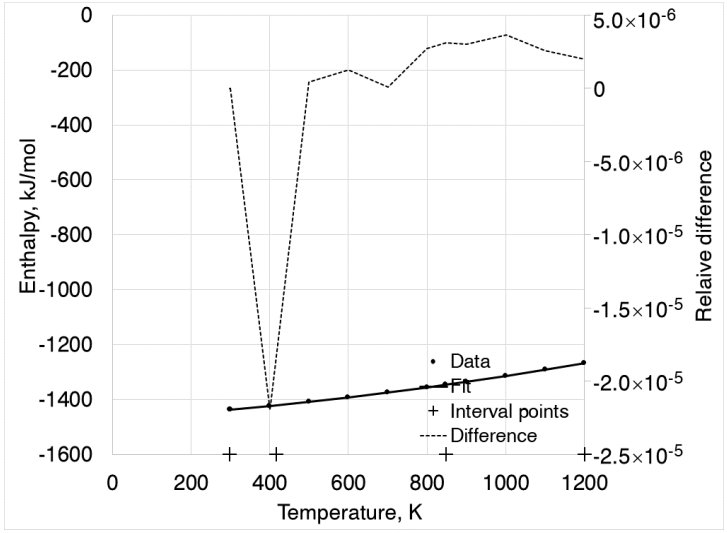


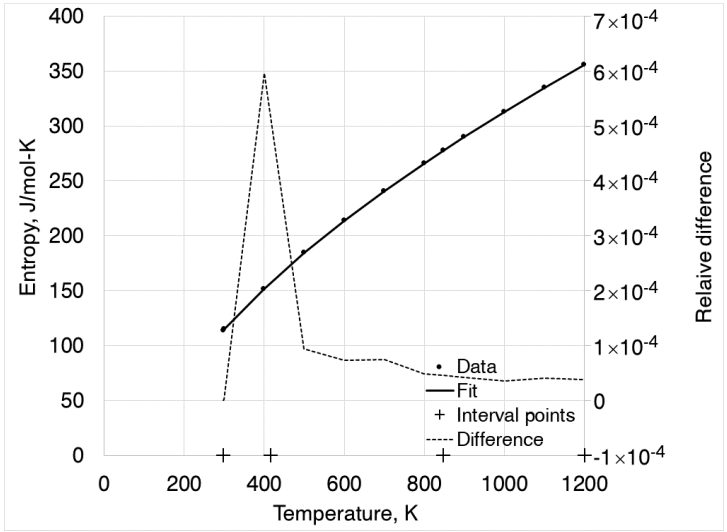


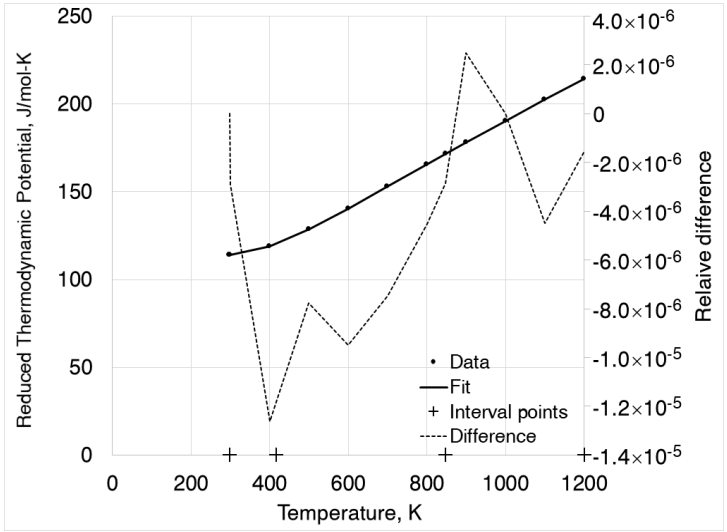
Lithium Sulfate, Alpha (Li_2SO_4) – Li2O4S1(cr)_01

MW = 109.94 g/mol – $\Delta_f H^o$(298.15 K) = -1436.367 kJ/mol – S^o(298.15 K) = 113.947 J/mol-K – p^o = 0.1 MPa



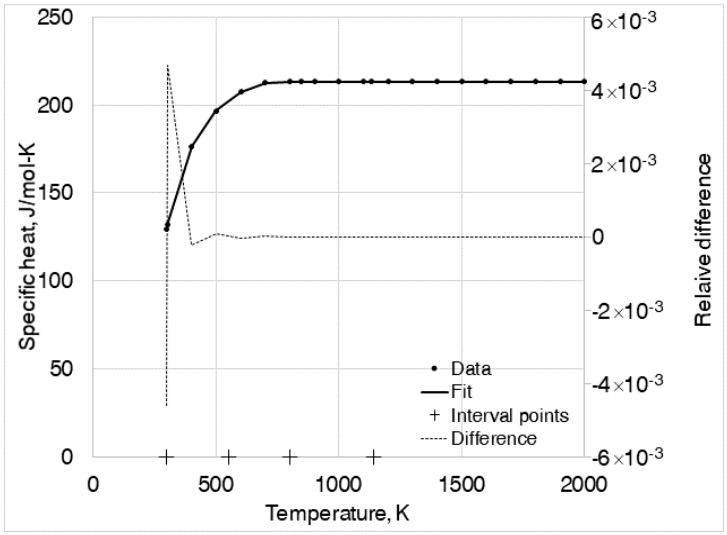


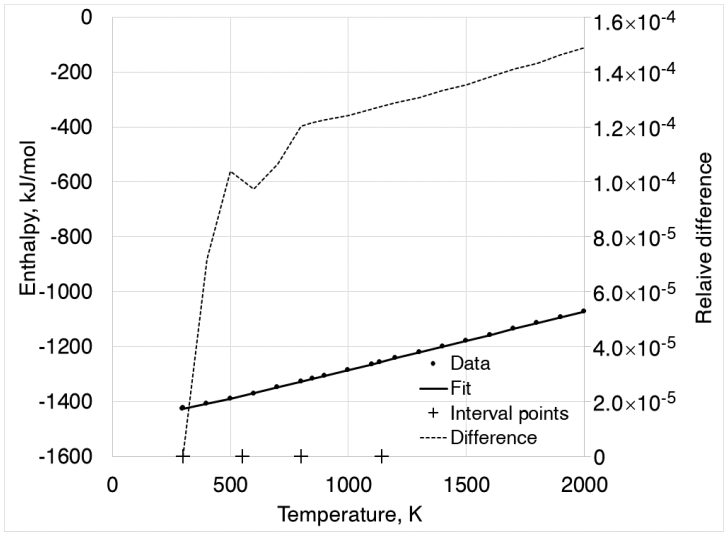


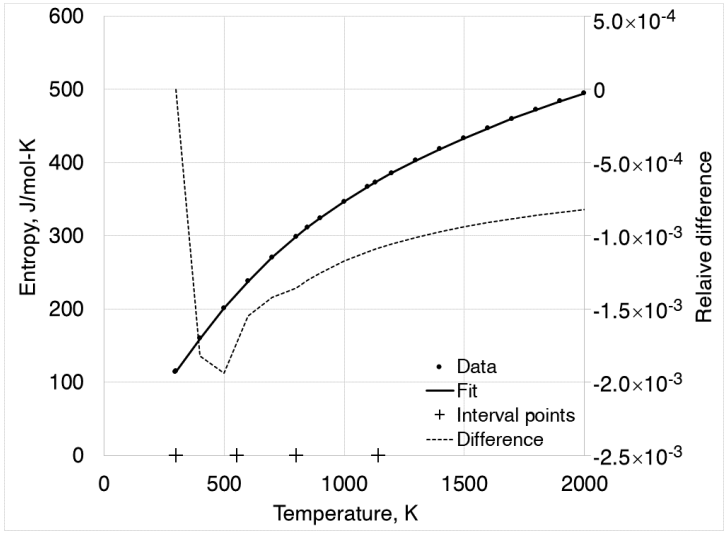


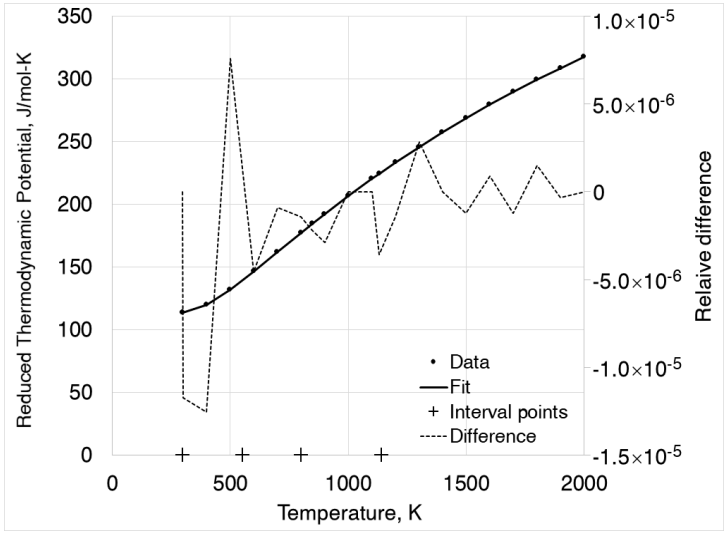
Lithium Sulfate, Beta (Li_2SO_4) – Li2O4S1(cr)_02

MW = 109.94 g/mol – $\Delta_f H^o$(298.15 K) = -1425.373 kJ/mol – S^o(298.15 K) = 113.802 J/mol-K – p^o = 0.1 MPa



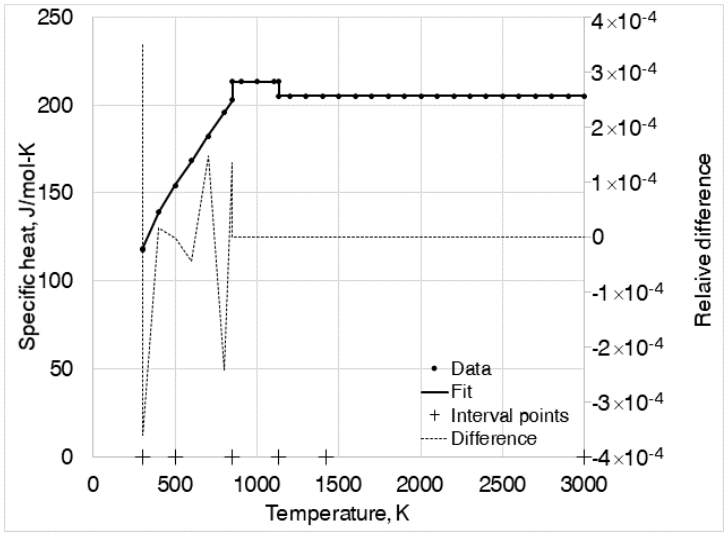


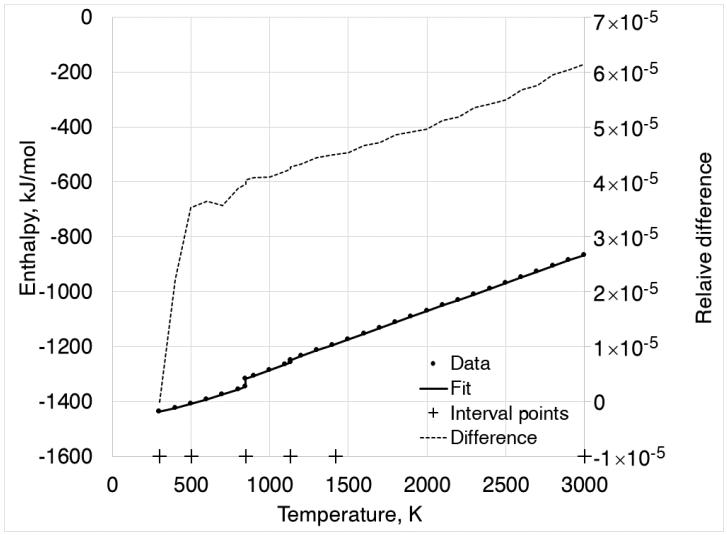


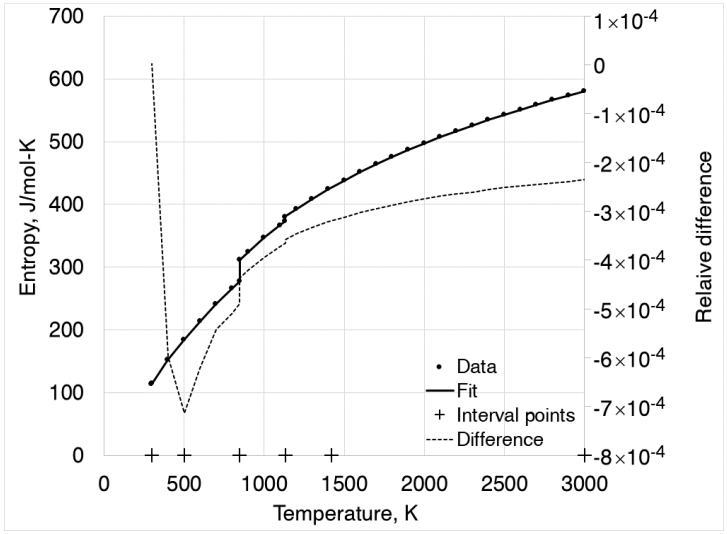


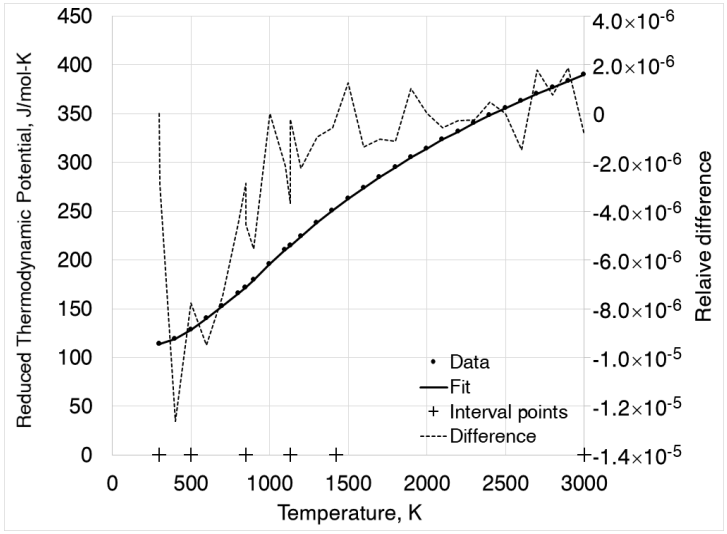
Lithium Sulfate (Li_2SO_4) – Li2O4S1(cr,l)

MW = 109.94 g/mol – $\Delta_f H^o$(298.15 K) = -1436.367 kJ/mol – S^o(298.15 K) = 113.947 J/mol-K – p^o = 0.1 MPa



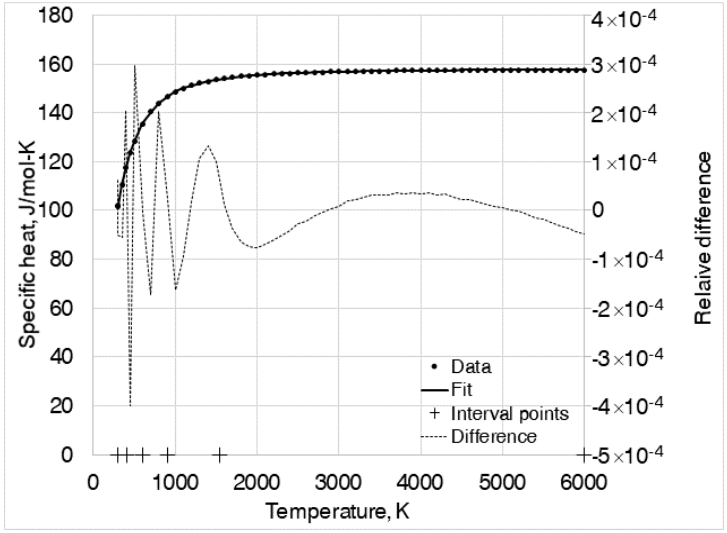


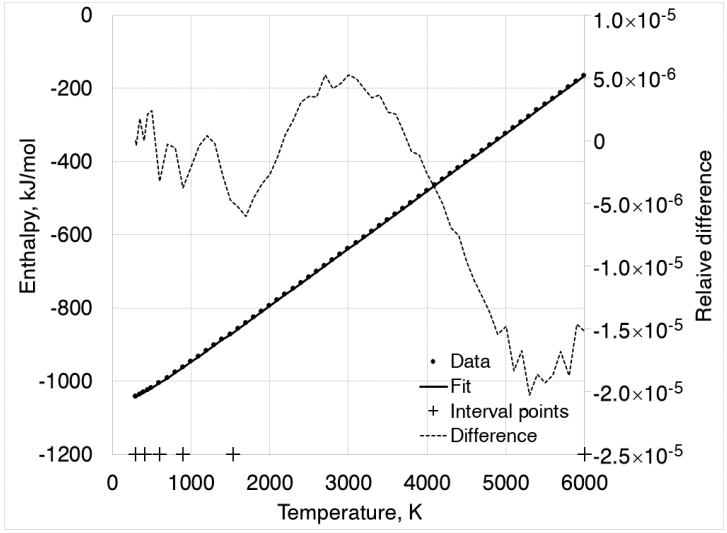


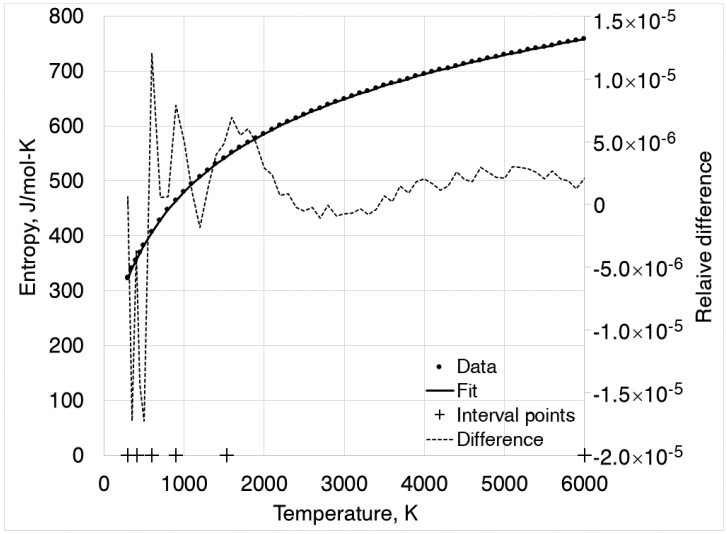


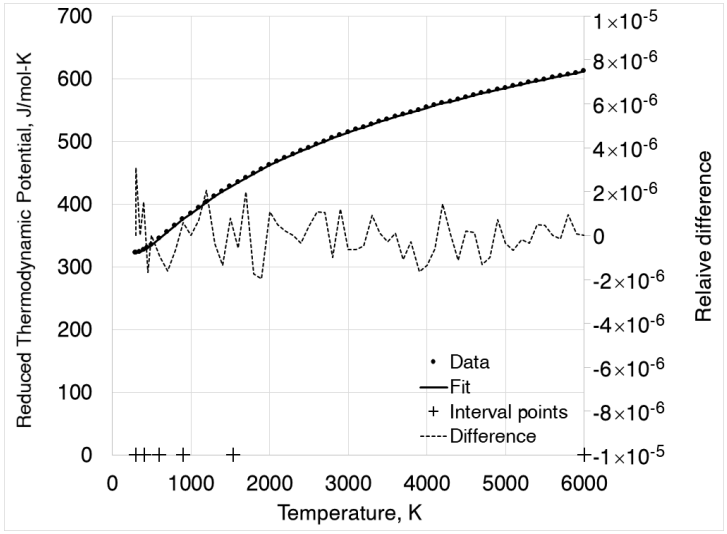
Lithium Sulfate (Li_2SO_4) – Li2O4S1(g)

MW = 109.94 g/mol – $\Delta_f H^o$(298.15 K) = -1041.816 kJ/mol – S^o(298.15 K) = 322.82 J/mol-K – p^o = 0.1 MPa



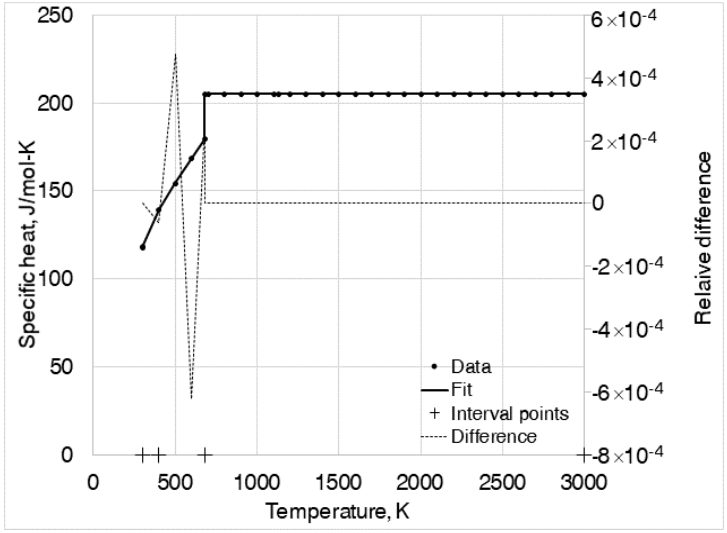


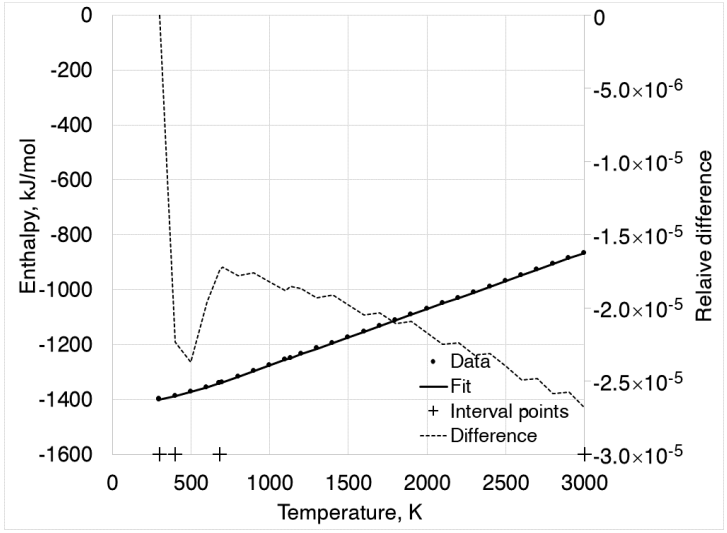


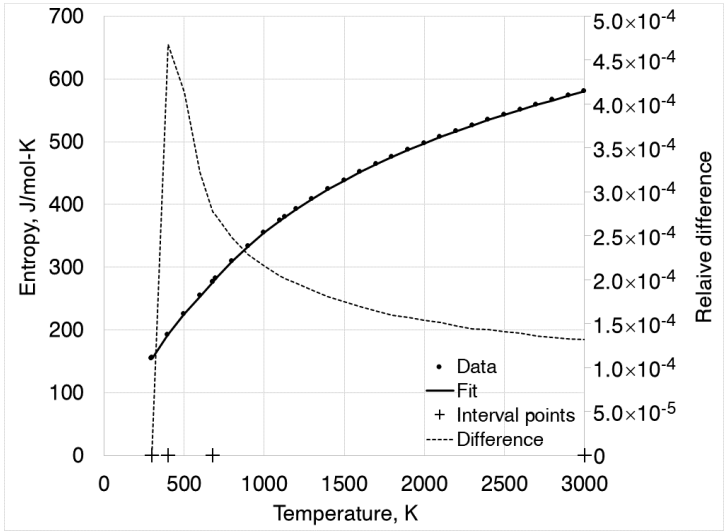


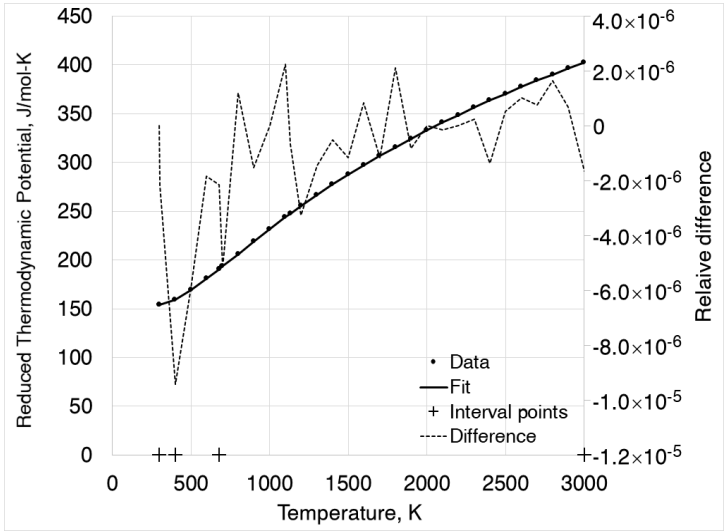
Lithium Sulfate (Li_2SO_4) – Li2O4S1(l)

MW = 109.94 g/mol – $\Delta_f H^o$(298.15 K) = -1399.296 kJ/mol – S^o(298.15 K) = 154.332 J/mol-K – p^o = 0.1 MPa



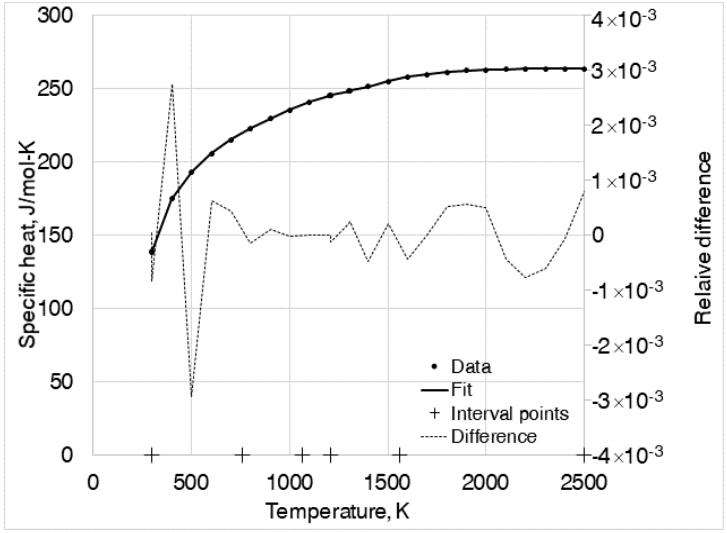


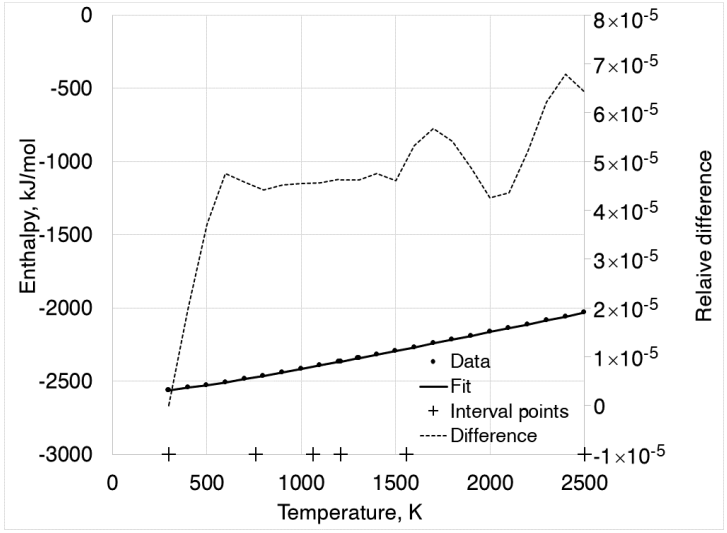


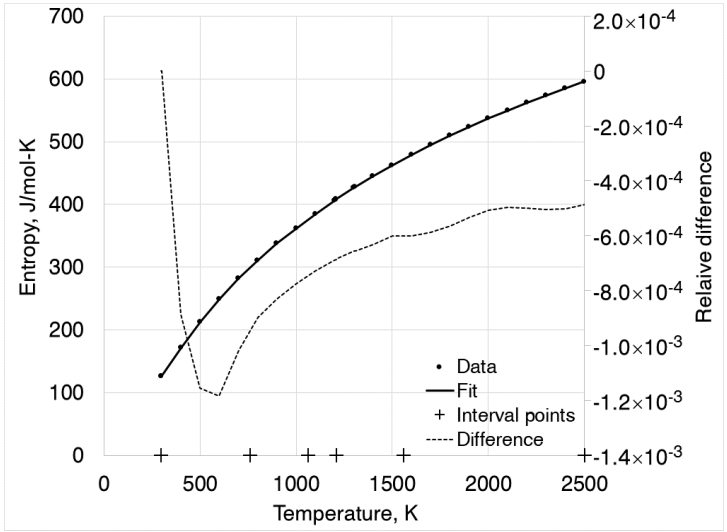


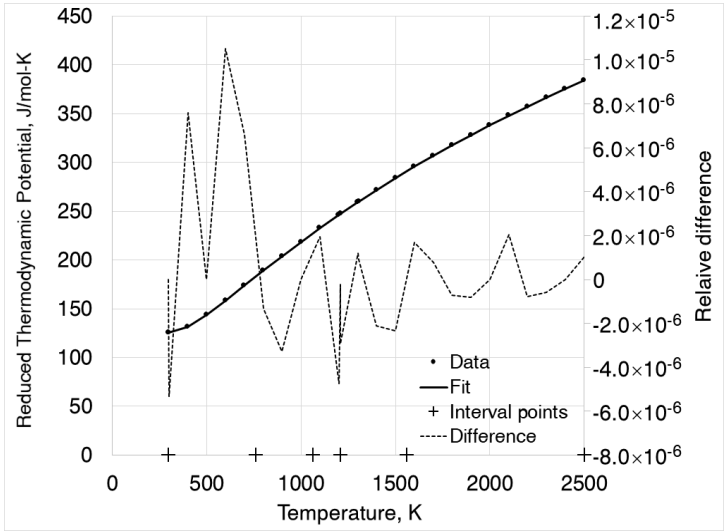
Lithium Silicate ($Li_2Si_2O_5$) – Li2O5Si2(cr)

MW = 150.05 g/mol – $\Delta_f H^o$(298.15 K) = -2560.901 kJ/mol – S^o(298.15 K) = 125.52 J/mol-K – p^o = 0.1 MPa



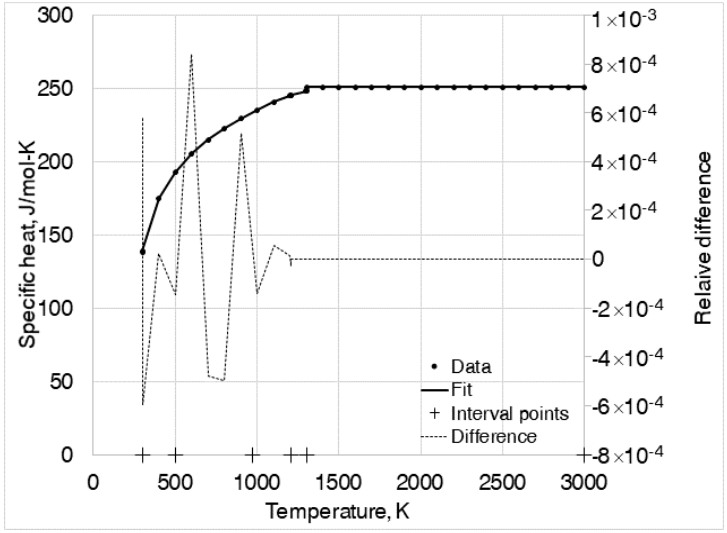


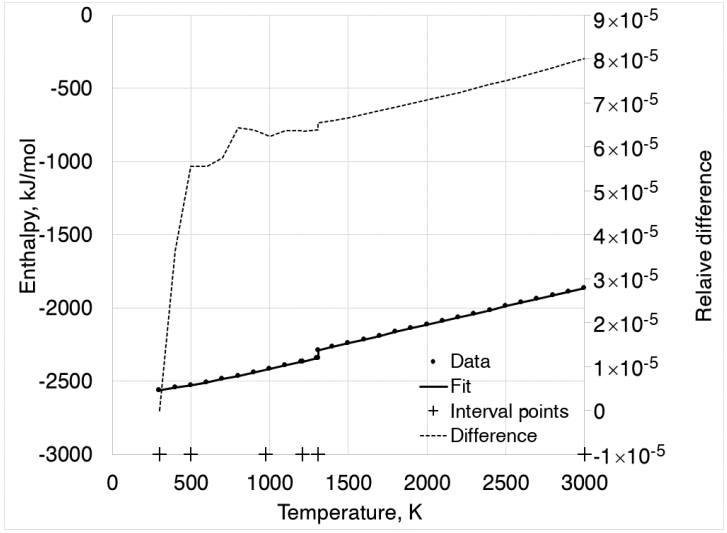


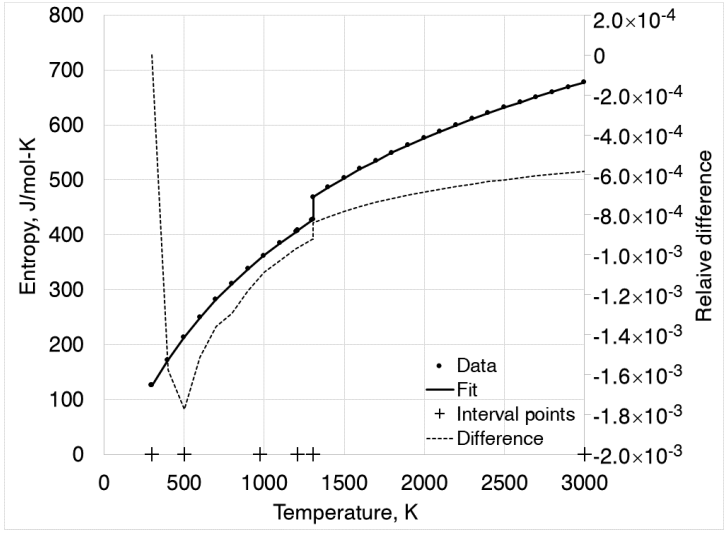


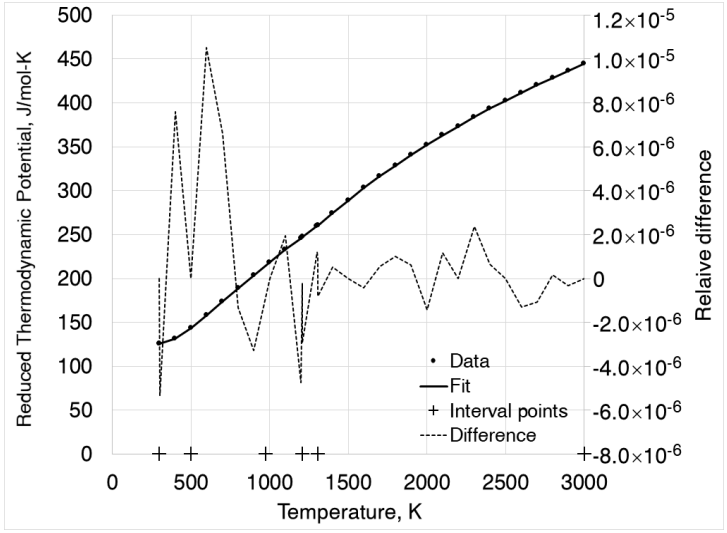
Lithium Silicate ($Li_2Si_2O_5$) – Li2O5Si2(cr,l)

MW = 150.05 g/mol – $\Delta_f H^o$(298.15 K) = -2560.901 kJ/mol – S^o(298.15 K) = 125.52 J/mol-K – p^o = 0.1 MPa



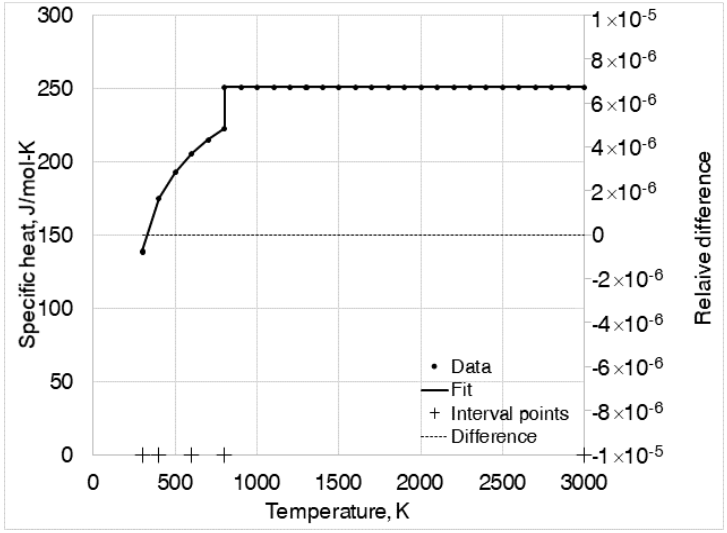


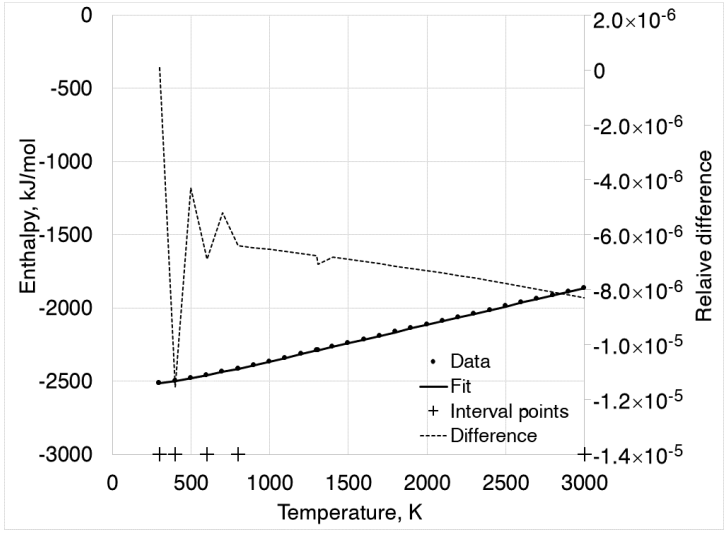


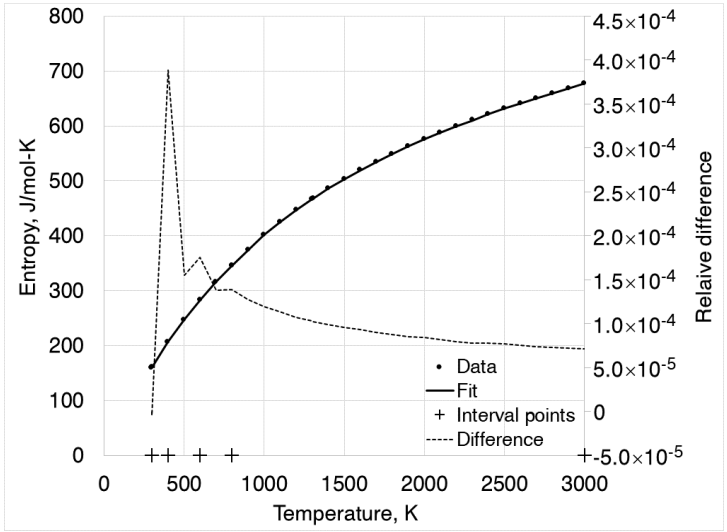


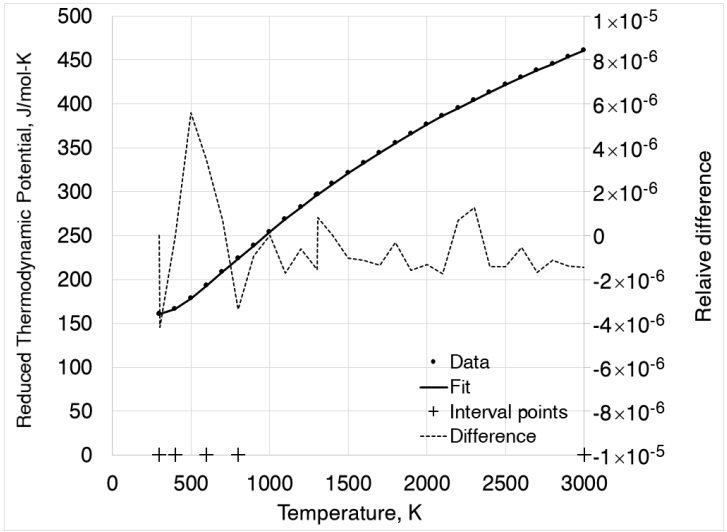
Lithium Silicate ($Li_2Si_2O_5$) – Li2O5Si2(l)

MW = 150.05 g/mol – $\Delta_f H^o$(298.15 K) = -2513.068 kJ/mol – S^o(298.15 K) = 160.243 J/mol-K – p^o = 0.1 MPa



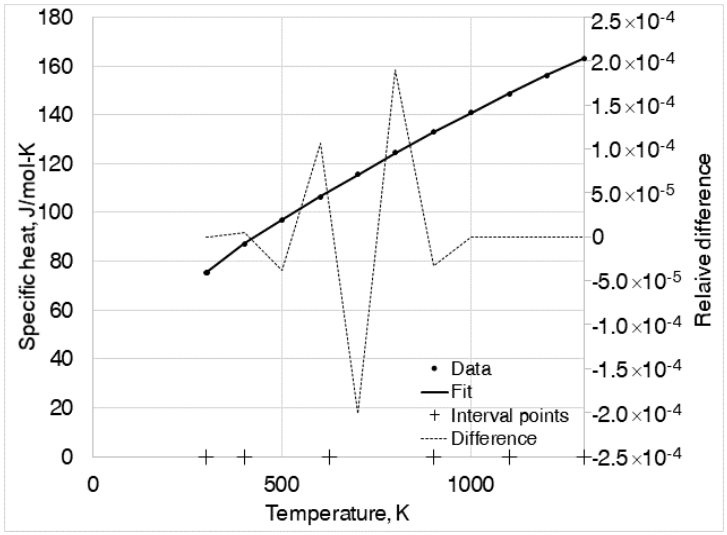


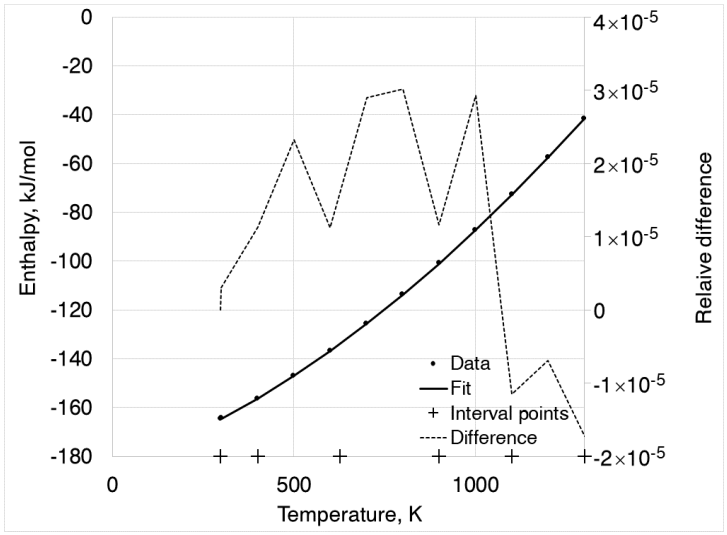


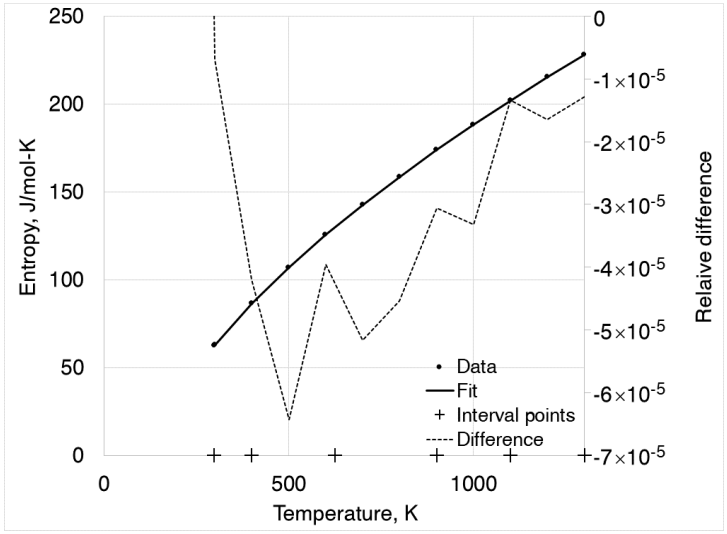


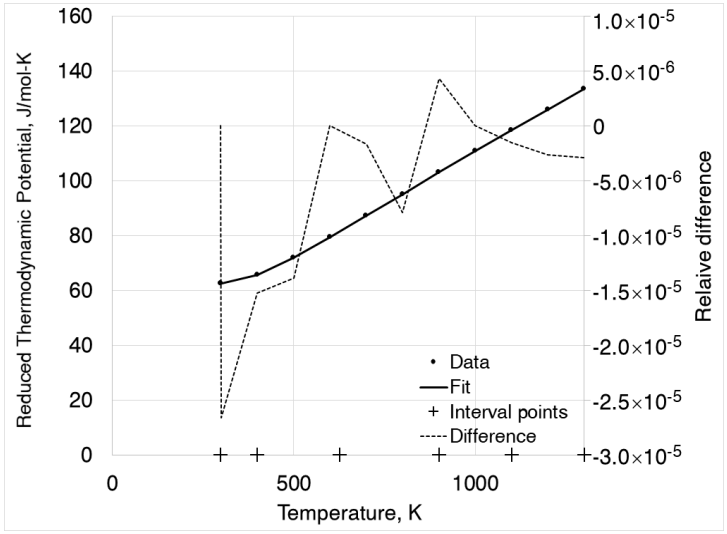
Lithium Nitride (Li_3N) – Li3N1(cr)

MW = 34.83 g/mol – $\Delta_f H^o$(298.15 K) = -164.557 kJ/mol – S^o(298.15 K) = 62.593 J/mol-K – p^o = 0.1 MPa



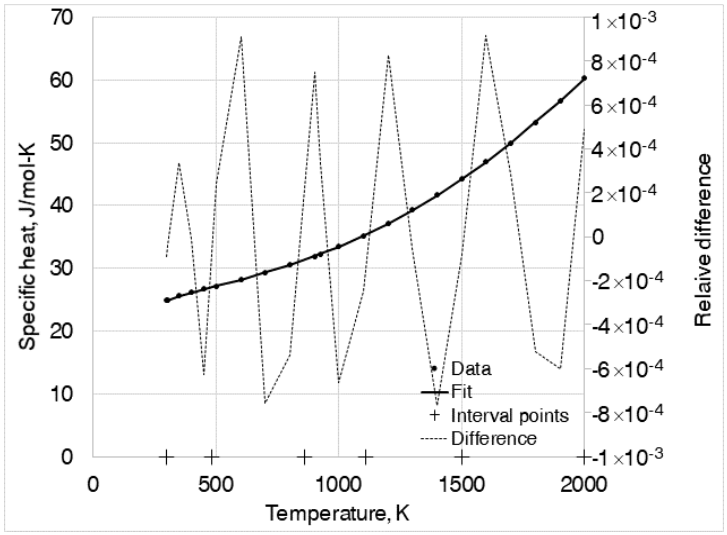


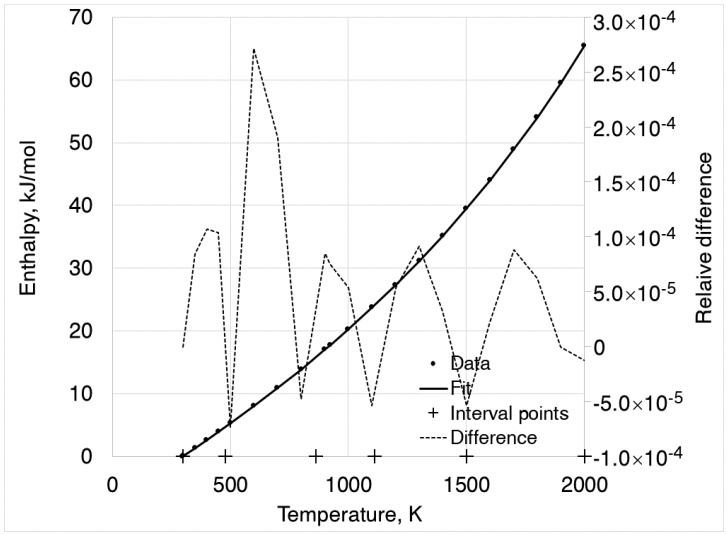


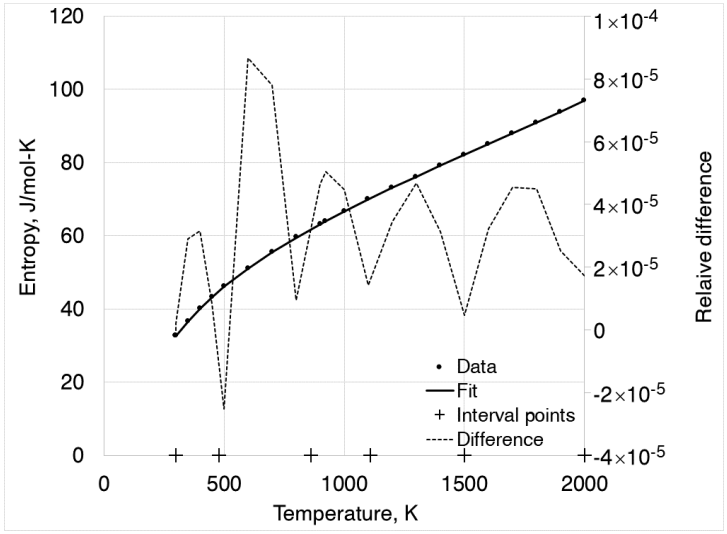


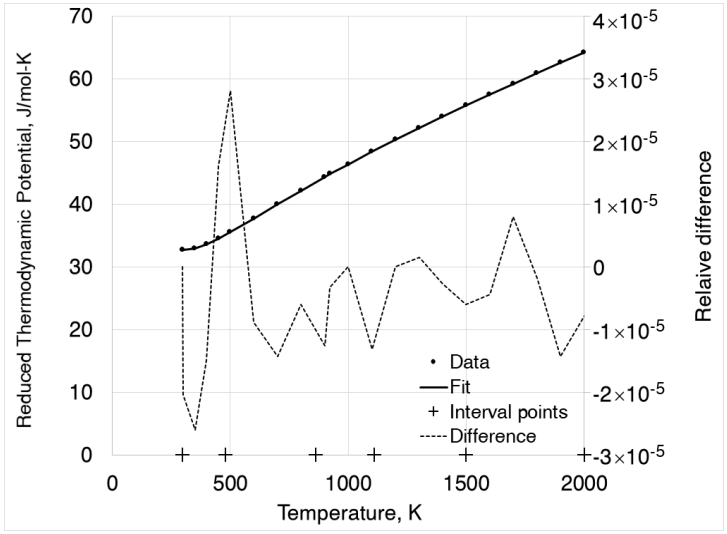
Magnesium (Mg) – Mg1(cr)

MW = 24.305 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 32.671 J/mol-K – p^o = 0.1 MPa



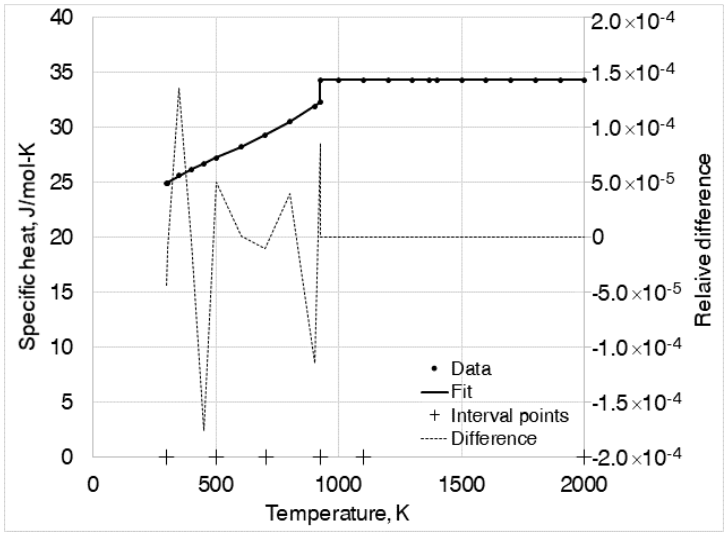


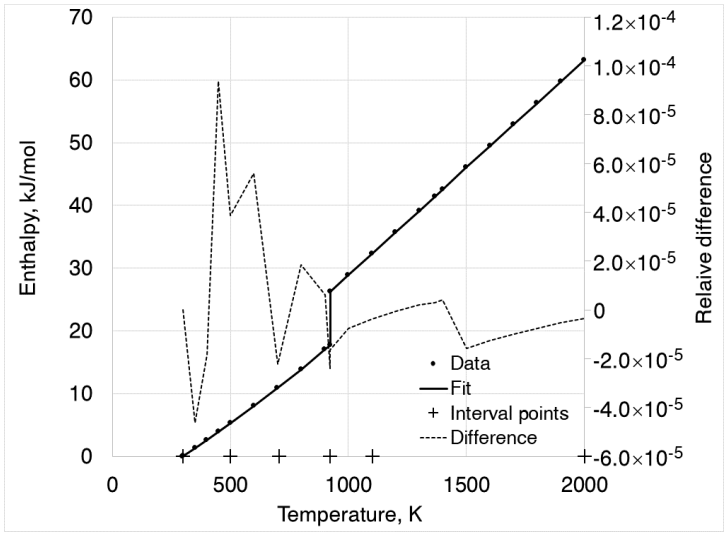


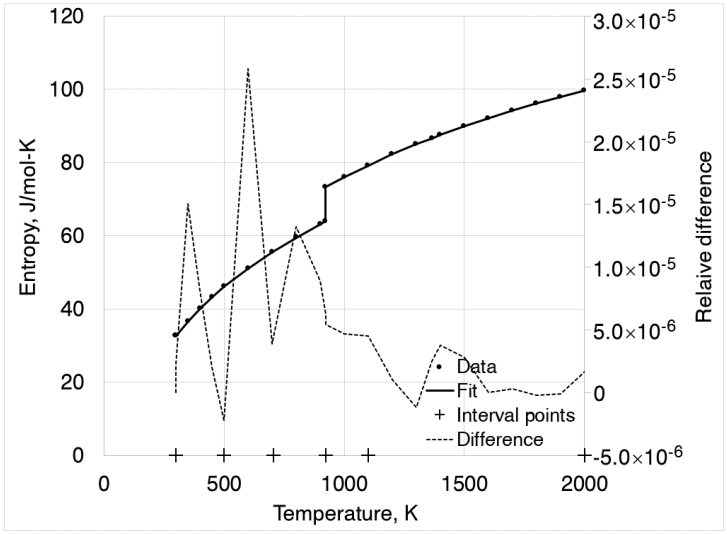


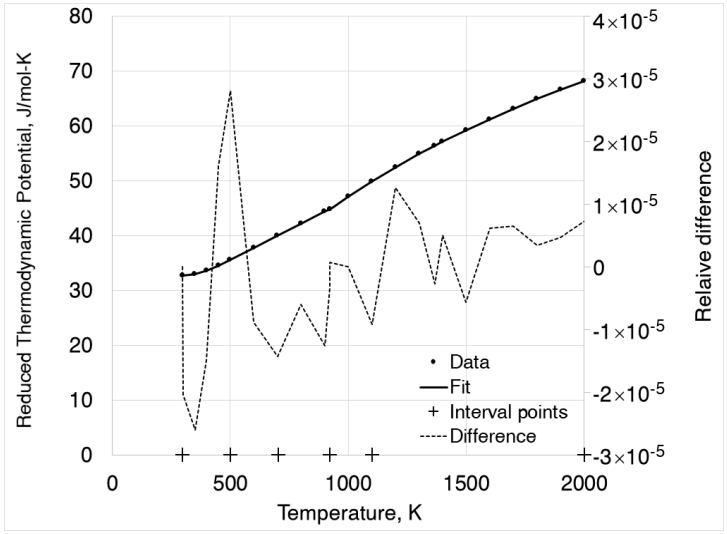
Magnesium (Mg) – Mg1(cr,l)

MW = 24.305 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 32.671 J/mol-K – p^o = 0.1 MPa



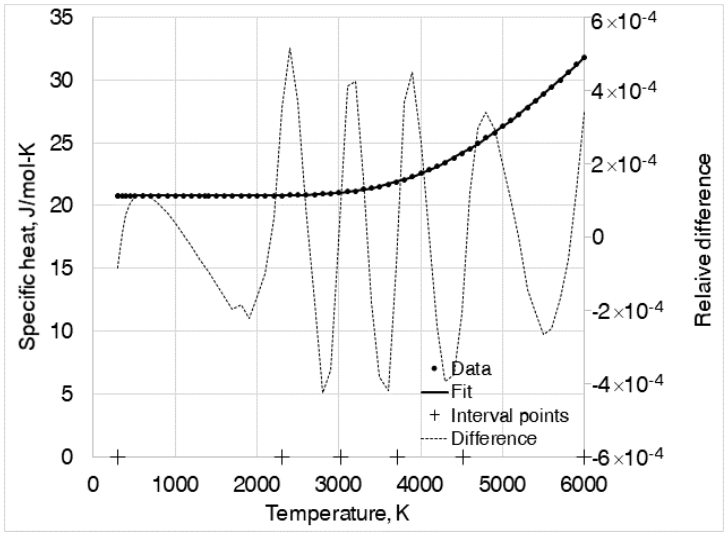


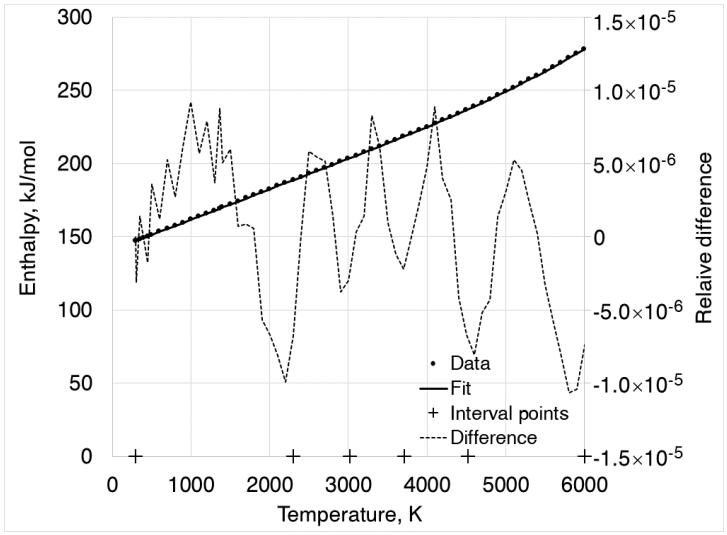


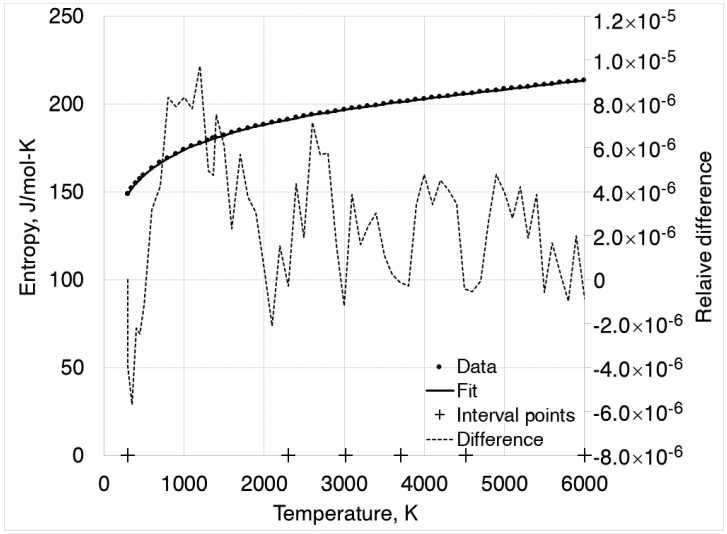


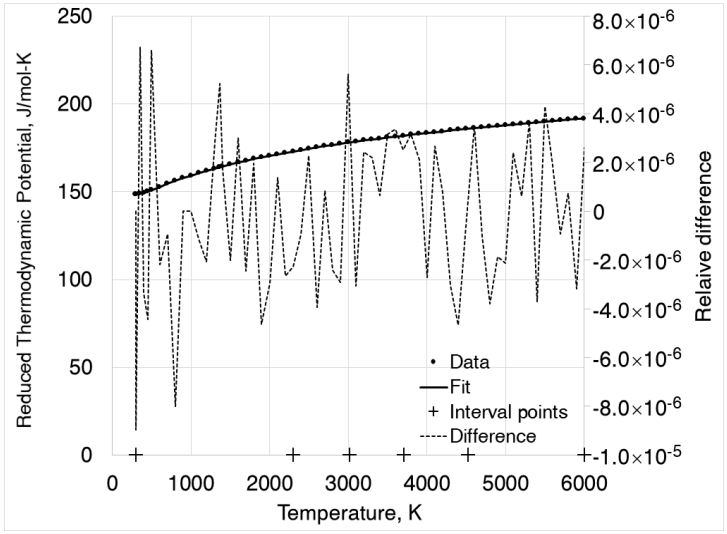
Magnesium (Mg) – Mg1(g)

MW = 24.305 g/mol – $\Delta_f H^o$(298.15 K) = 147.1 kJ/mol – S^o(298.15 K) = 148.648 J/mol-K – p^o = 0.1 MPa



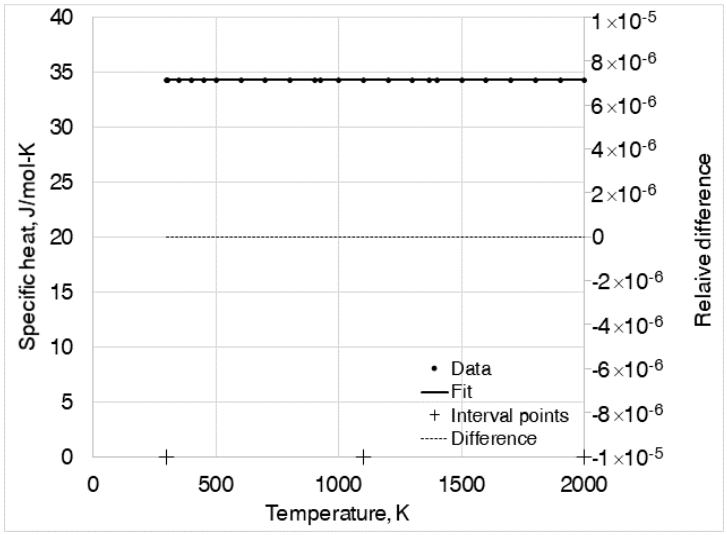


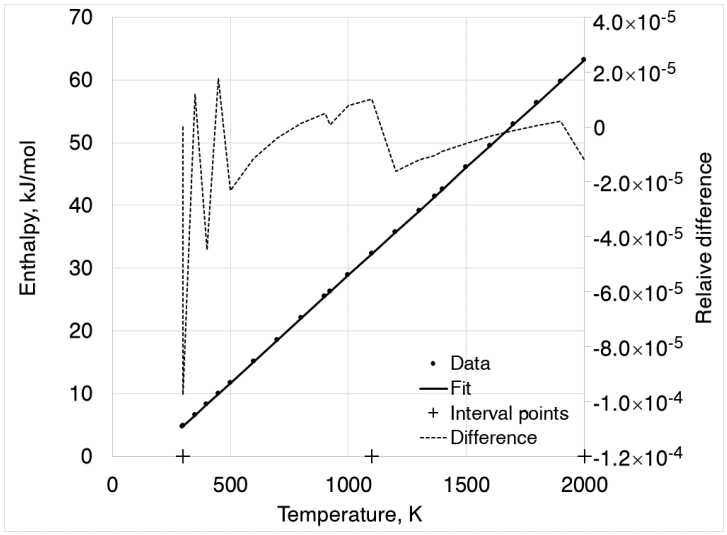


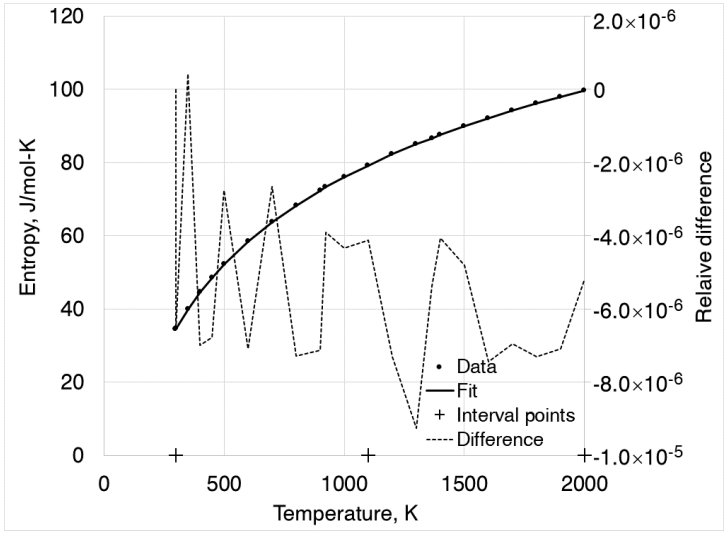


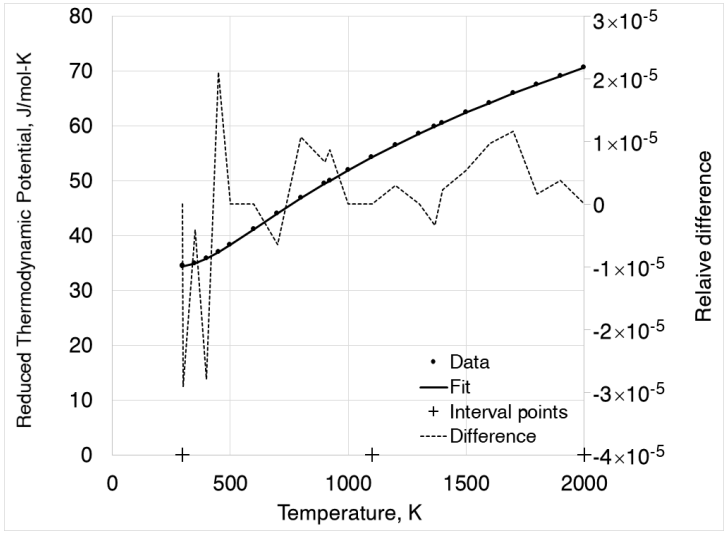
Magnesium (Mg) – Mg1(l)

MW = 24.305 g/mol – $\Delta_f H^o$(298.15 K) = 4.79 kJ/mol – S^o(298.15 K) = 34.464 J/mol-K – p^o = 0.1 MPa



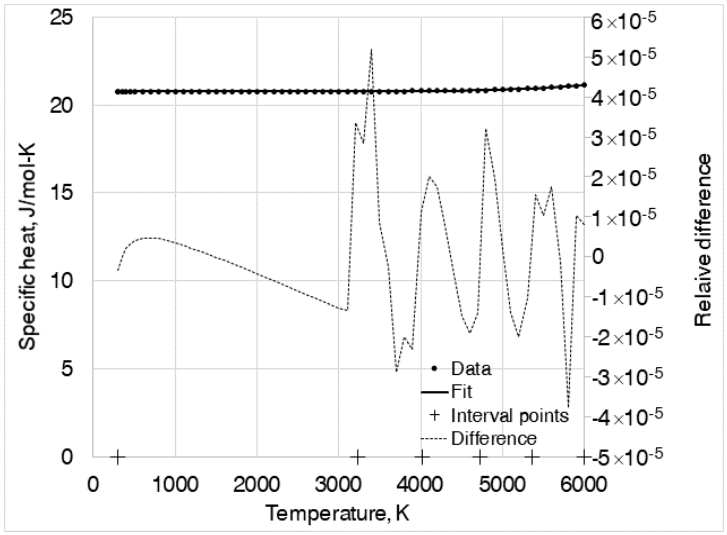


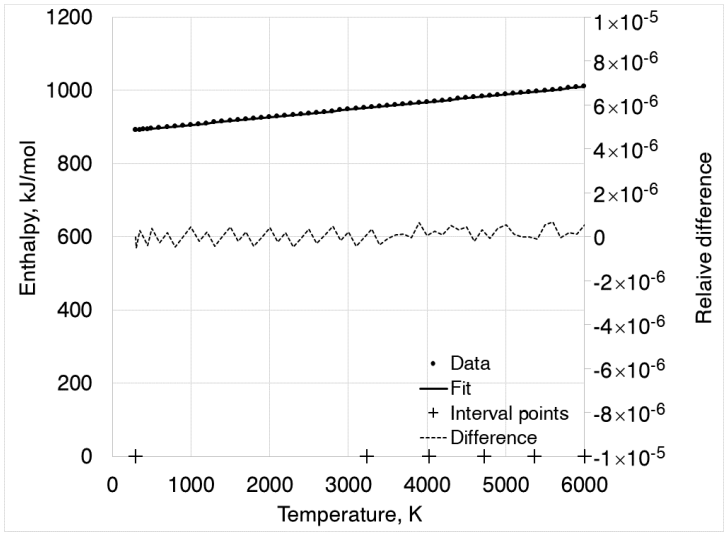


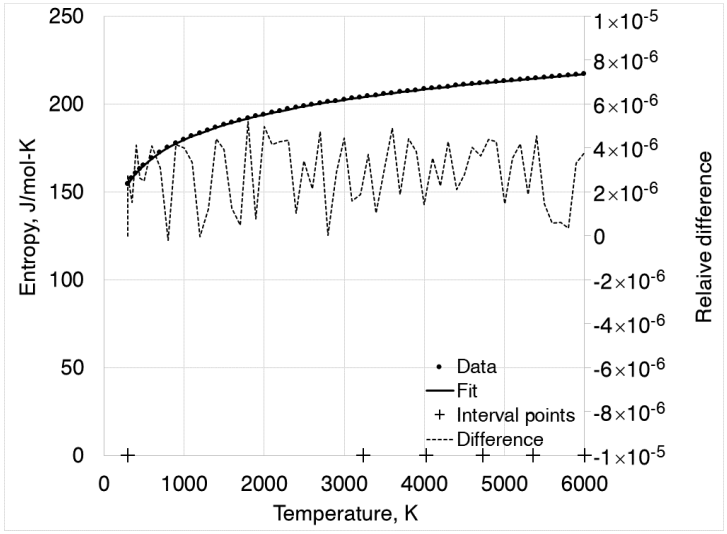


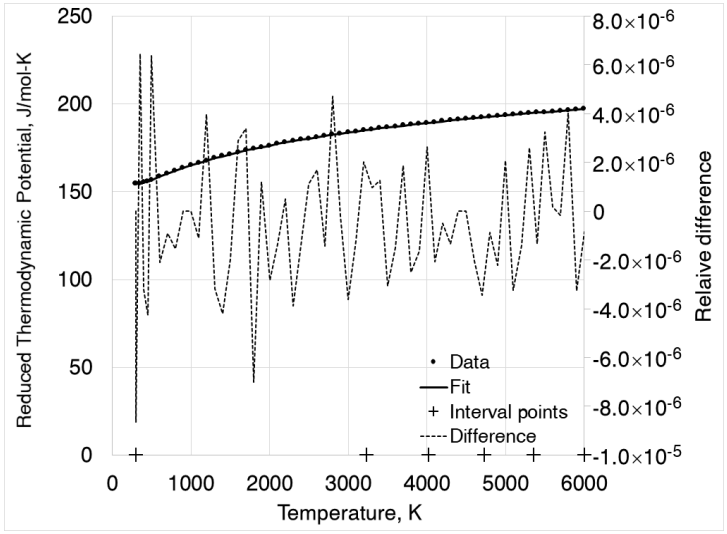
Magnesium, Ion (Mg^+) – Mg1+(g)

MW = 24.304 g/mol – $\Delta_f H^o$(298.15 K) = 891.047 kJ/mol – S^o(298.15 K) = 154.41 J/mol-K – p^o = 0.1 MPa



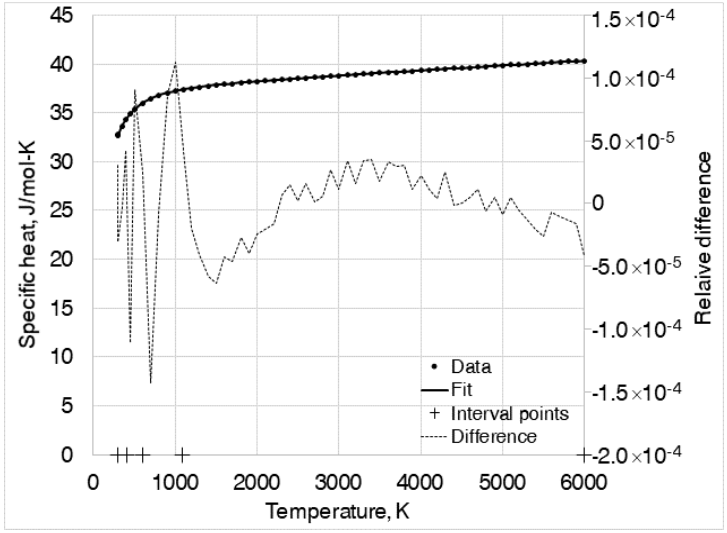


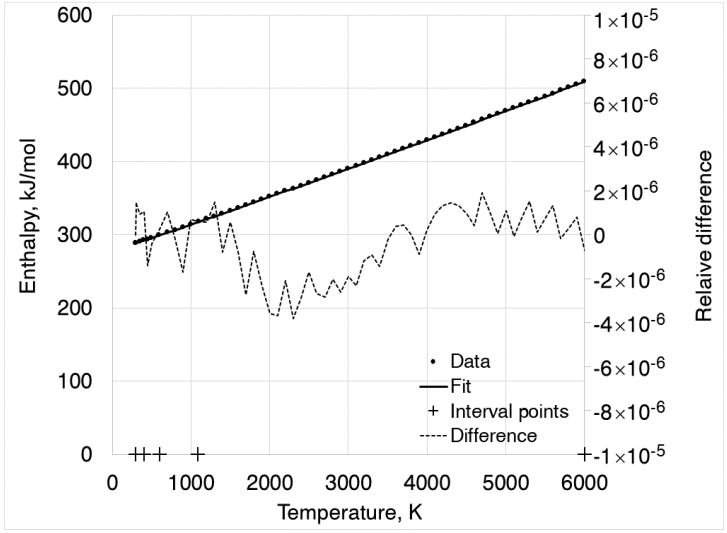


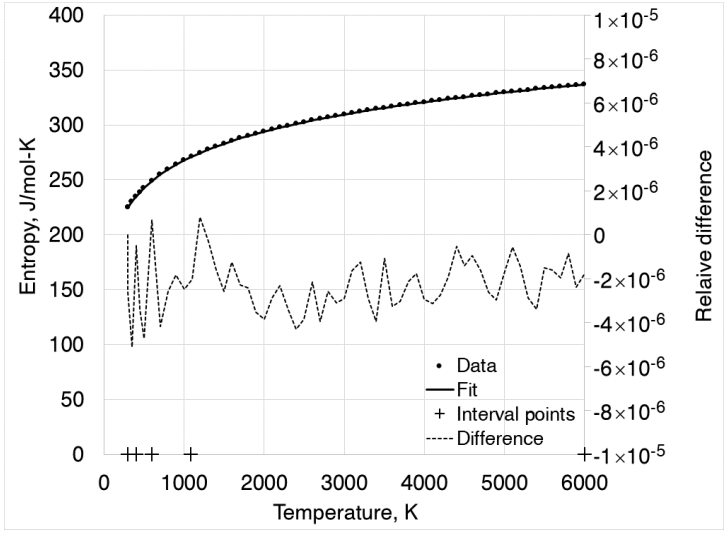


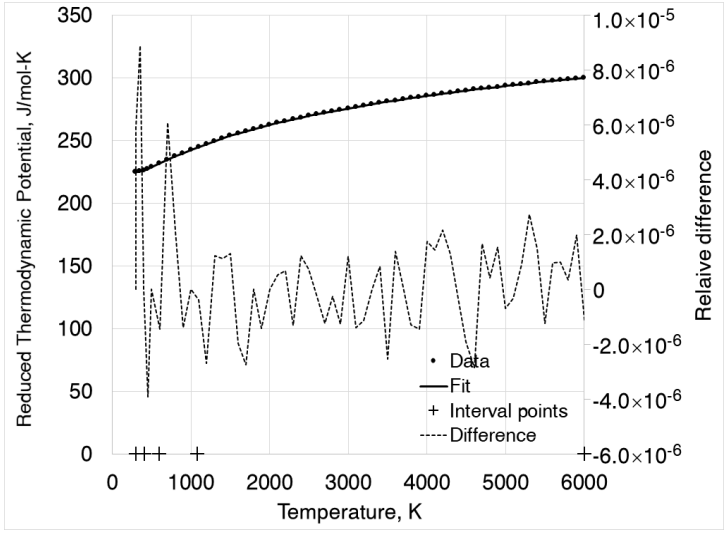
Magnesium Nitride (MgN) – Mg1N1(g)

MW = 38.312 g/mol – $\Delta_f H^o$(298.15 K) = 288.696 kJ/mol – S^o(298.15 K) = 224.835 J/mol-K – p^o = 0.1 MPa



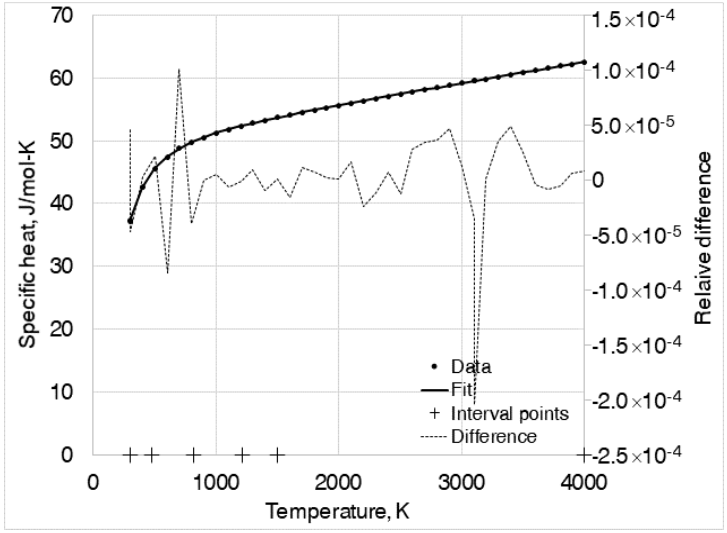


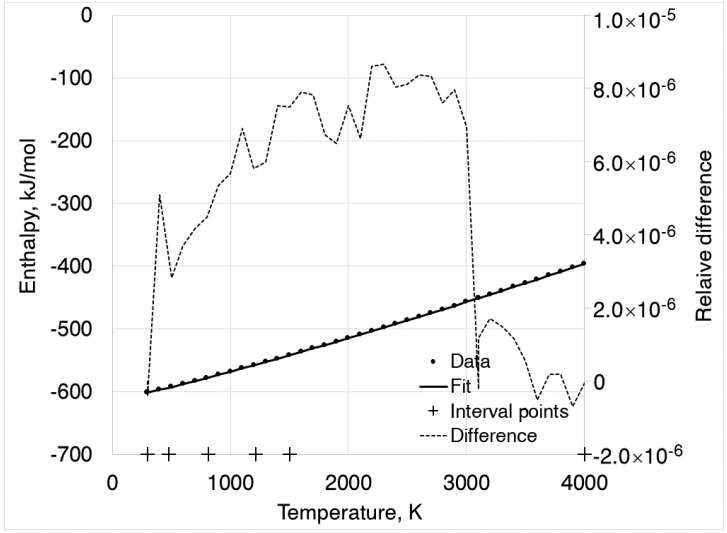


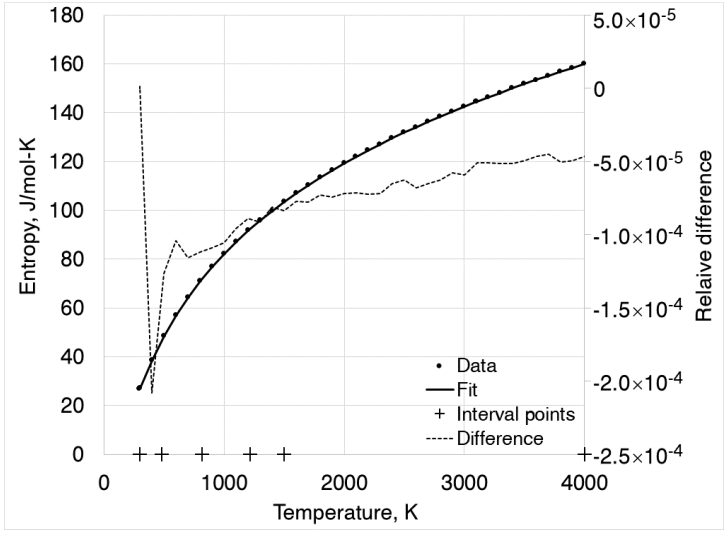


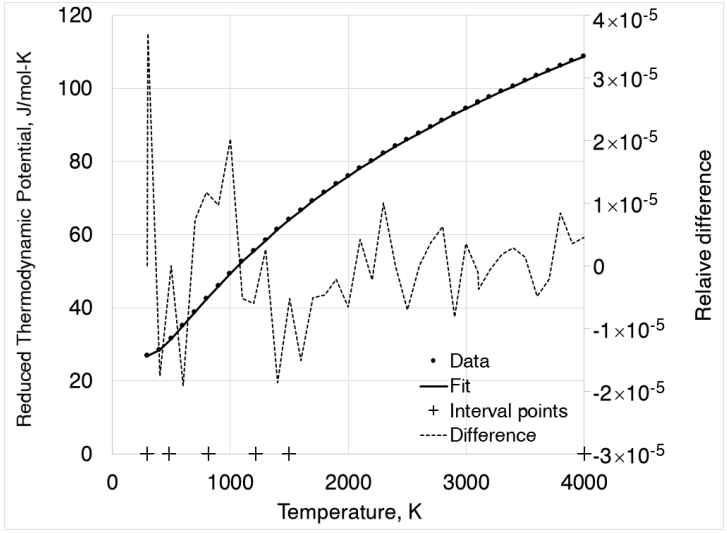
Magnesium Oxide (MgO) – Mg1O1(cr)

MW = 40.304 g/mol – $\Delta_f H^o$(298.15 K) = -601.241 kJ/mol – S^o(298.15 K) = 26.924 J/mol-K – p^o = 0.1 MPa



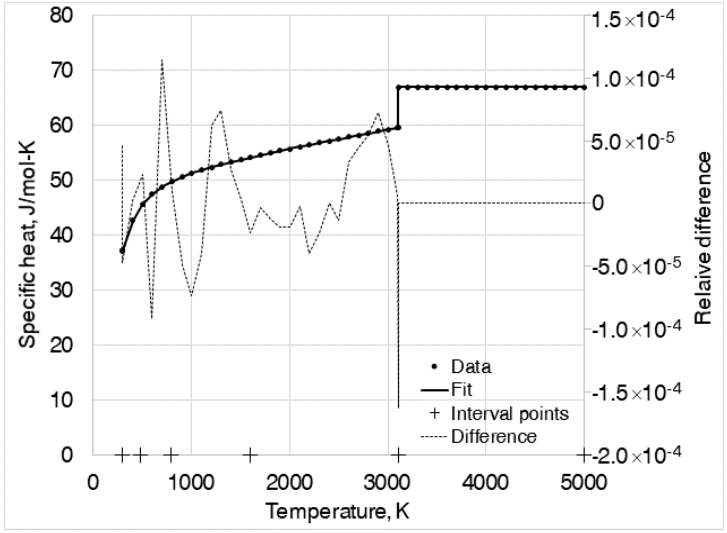


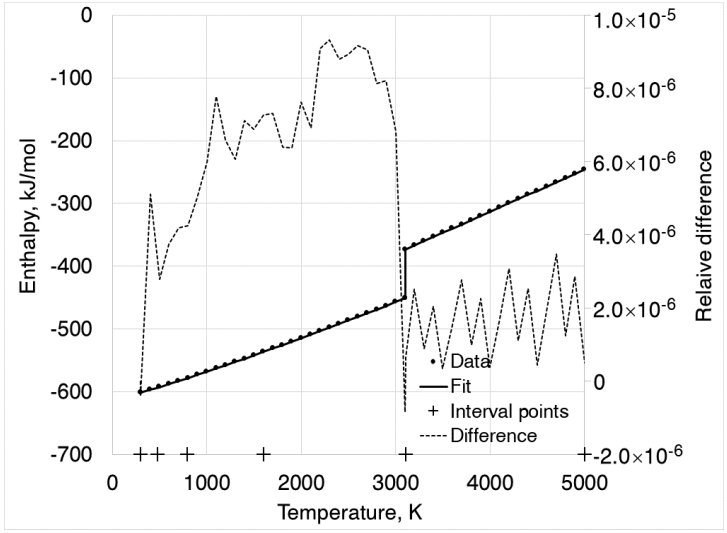


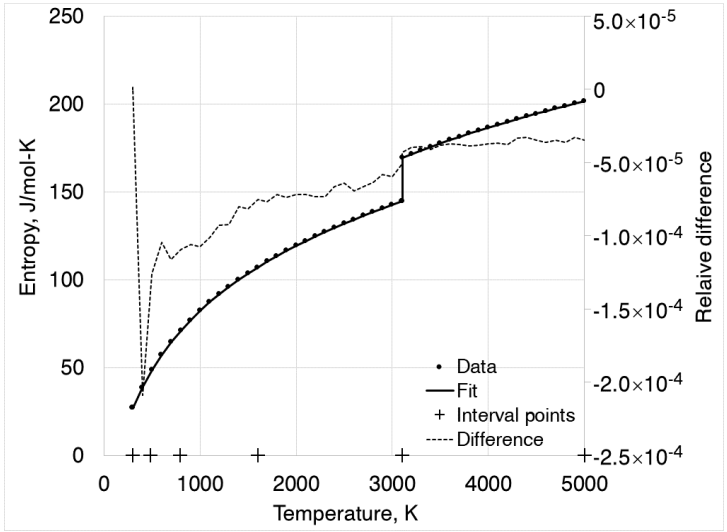


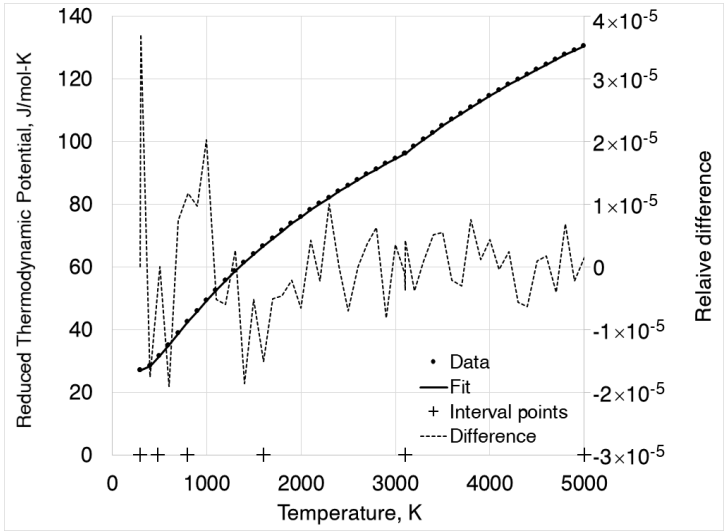
Magnesium Oxide (MgO) – Mg1O1(cr,l)

MW = 40.304 g/mol – $\Delta_f H^o$(298.15 K) = -601.241 kJ/mol – S^o(298.15 K) = 26.924 J/mol-K – p^o = 0.1 MPa



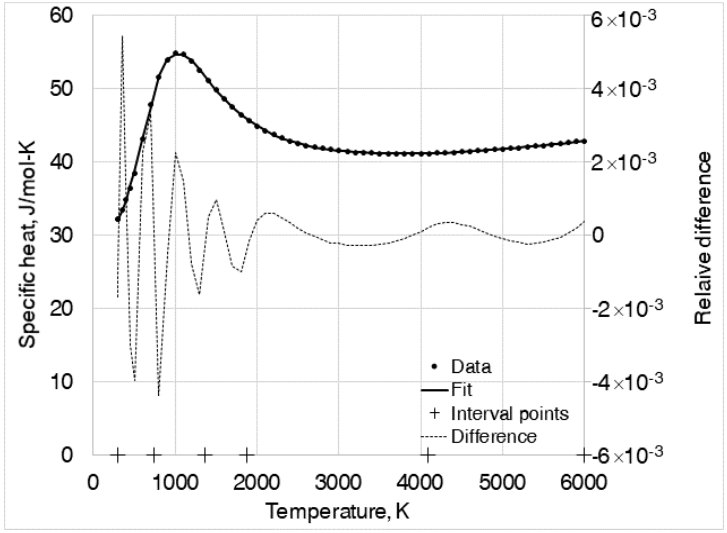


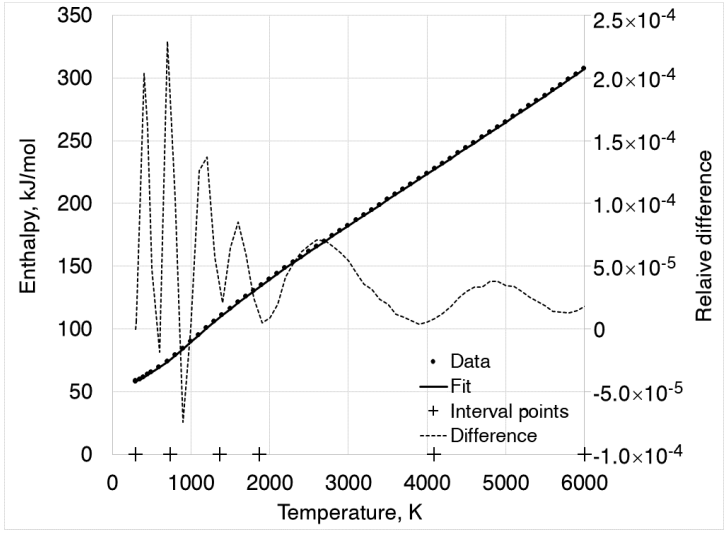


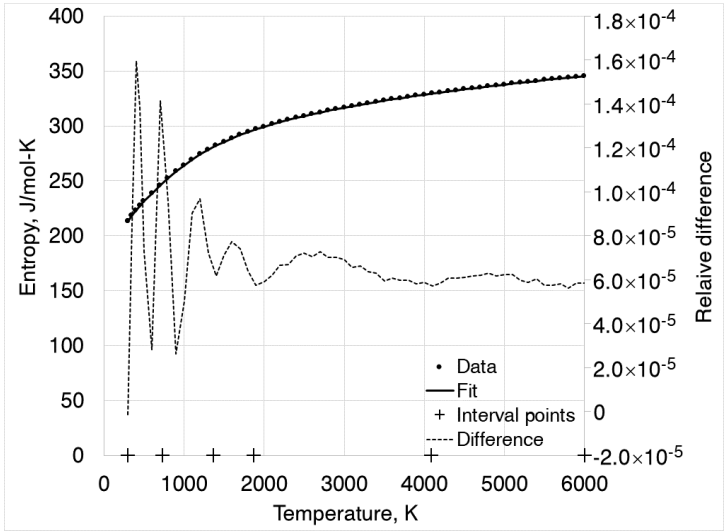


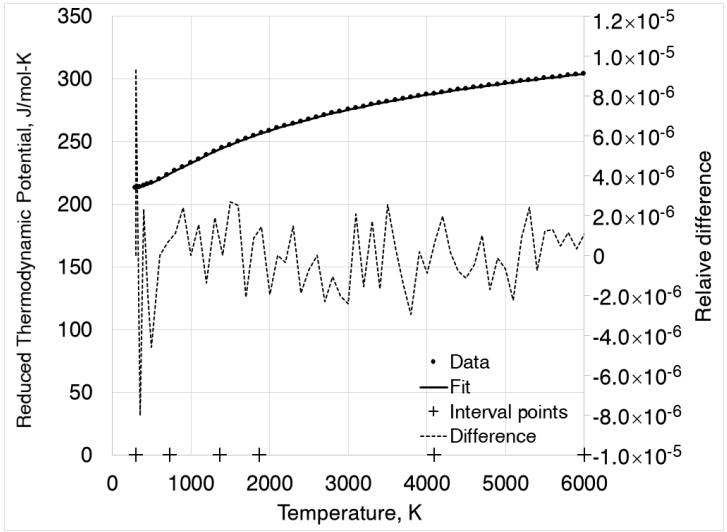
Magnesium Oxide (MgO) – Mg1O1(g)

MW = 40.304 g/mol – $\Delta_f H^o$(298.15 K) = 58.158 kJ/mol – S^o(298.15 K) = 213.269 J/mol-K – p^o = 0.1 MPa



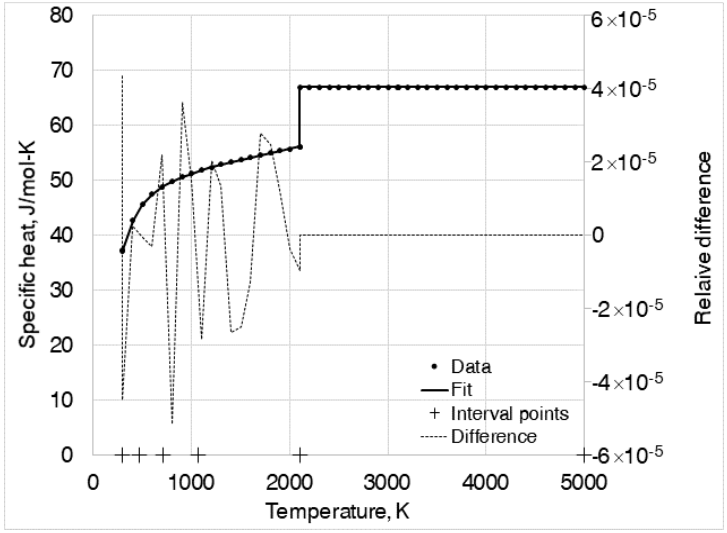


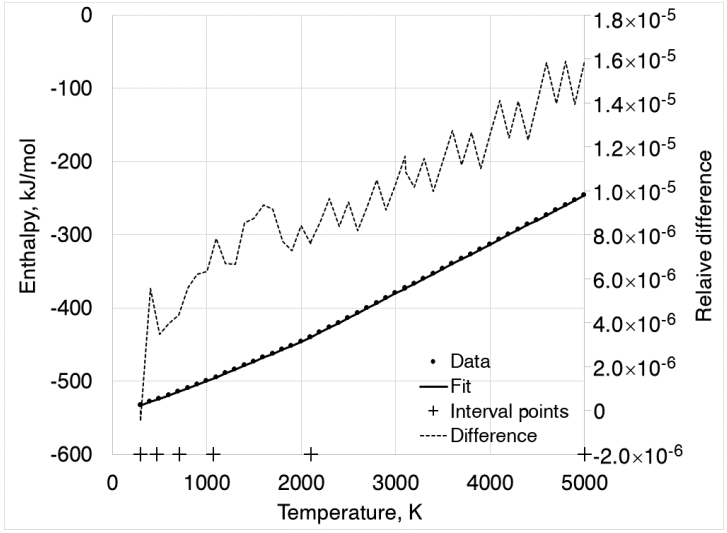


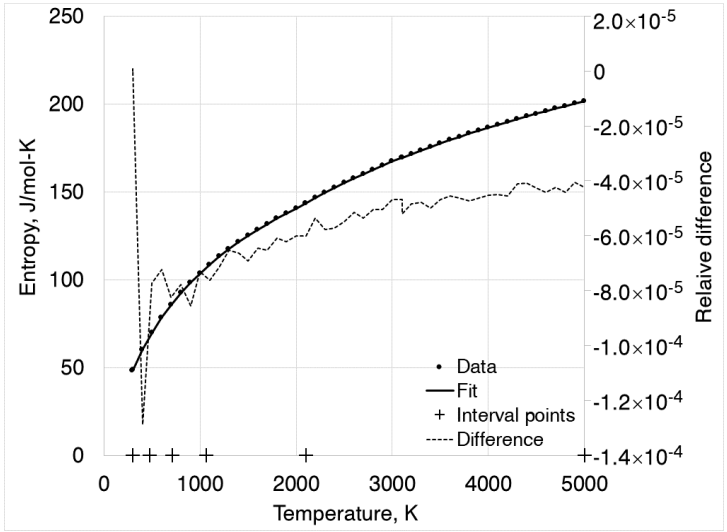


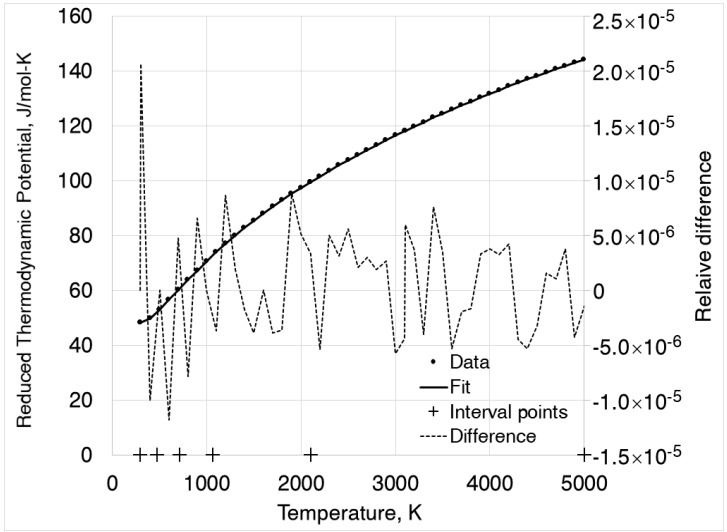
Magnesium Oxide (MgO) – Mg1O1(l)

MW = 40.304 g/mol – $\Delta_f H^o$(298.15 K) = -532.609 kJ/mol – S^o(298.15 K) = 48.366 J/mol-K – p^o = 0.1 MPa



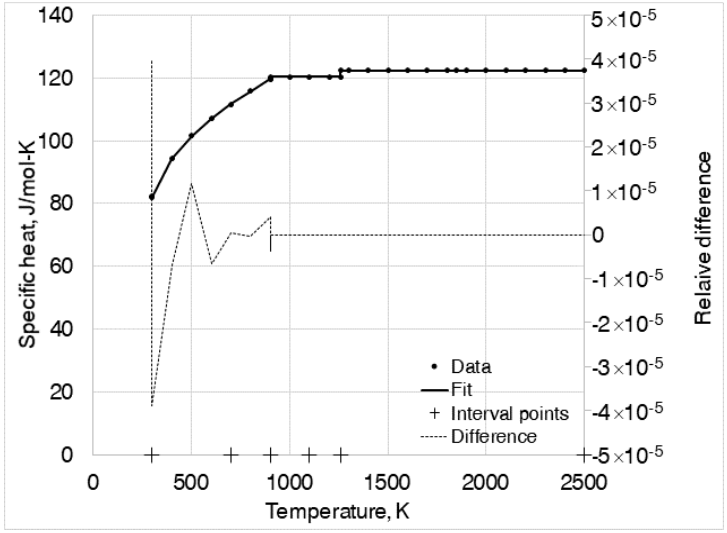


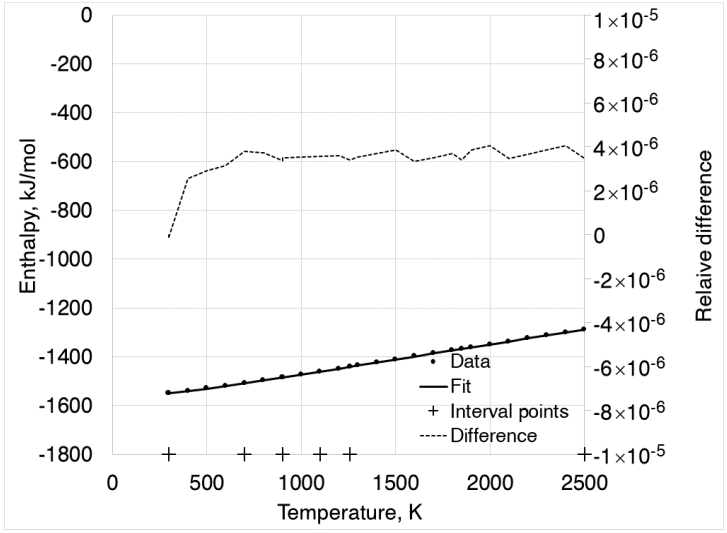


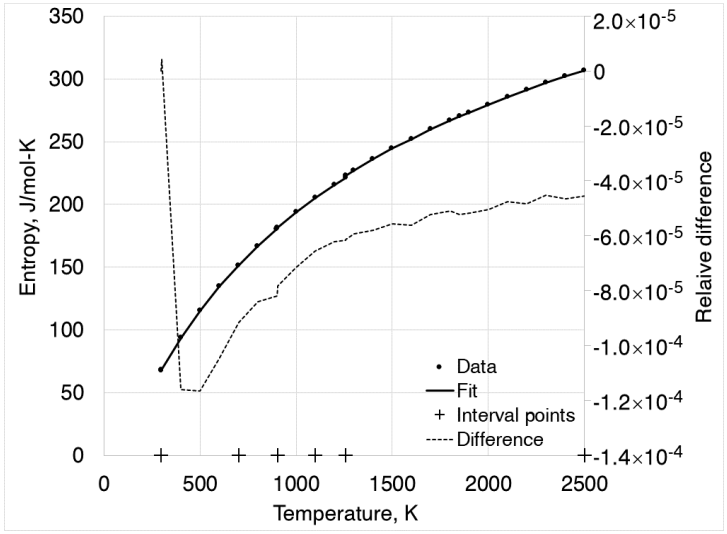


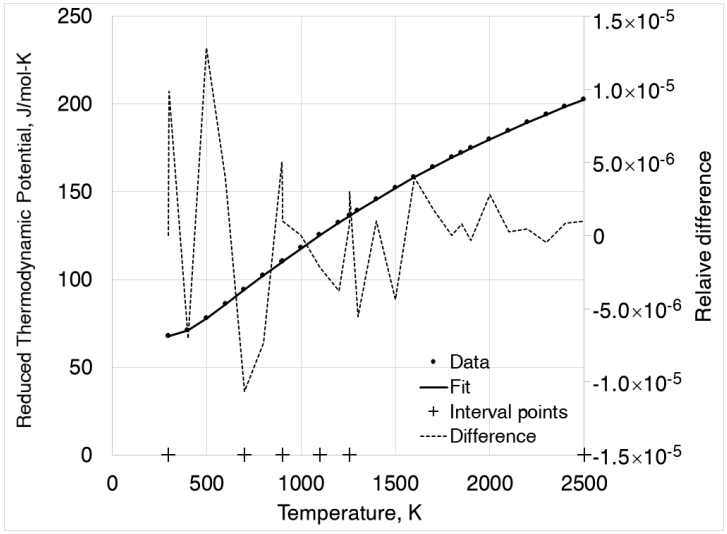
Magnesium Silicate ($MgSiO_3$) – Mg1O3Si1(cr)

MW = 100.389 g/mol – $\Delta_f H^o$(298.15 K) = -1548.917 kJ/mol – S^o(298.15 K) = 67.768 J/mol-K – p^o = 0.1 MPa



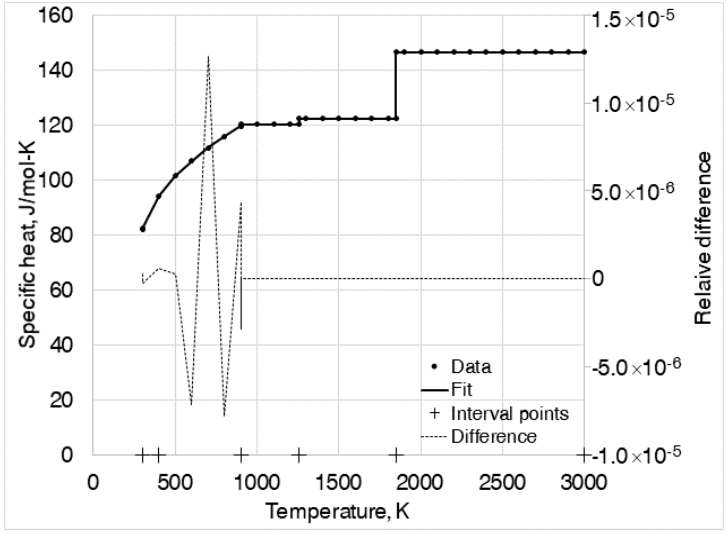


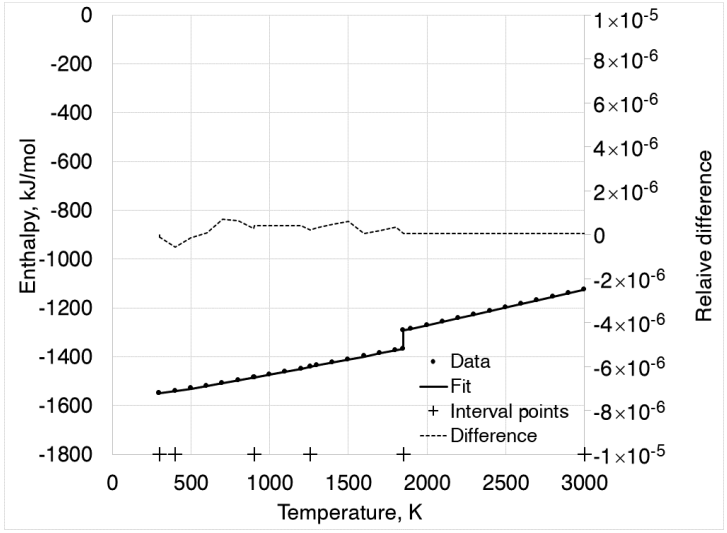


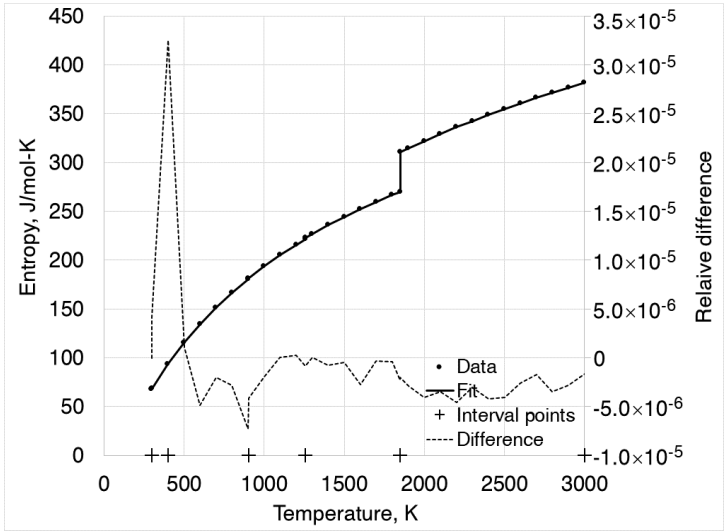


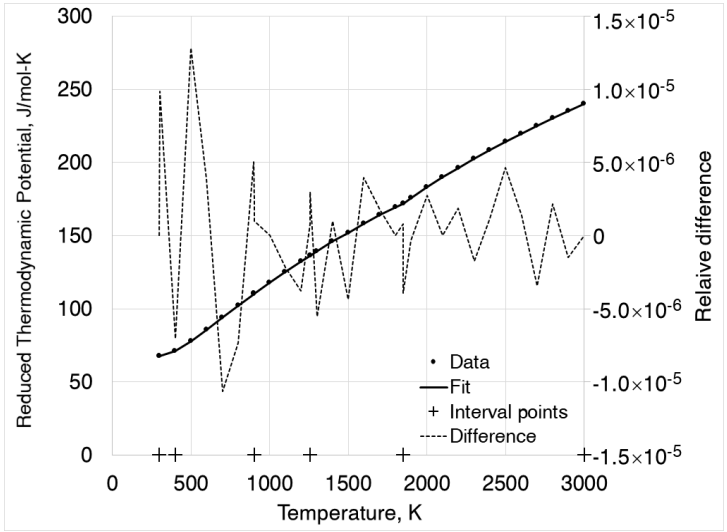
Magnesium Silicate ($MgSiO_3$) – Mg1O3Si1(cr,l)

MW = 100.389 g/mol – $\Delta_f H^o$(298.15 K) = -1548.917 kJ/mol – S^o(298.15 K) = 67.768 J/mol-K – p^o = 0.1 MPa



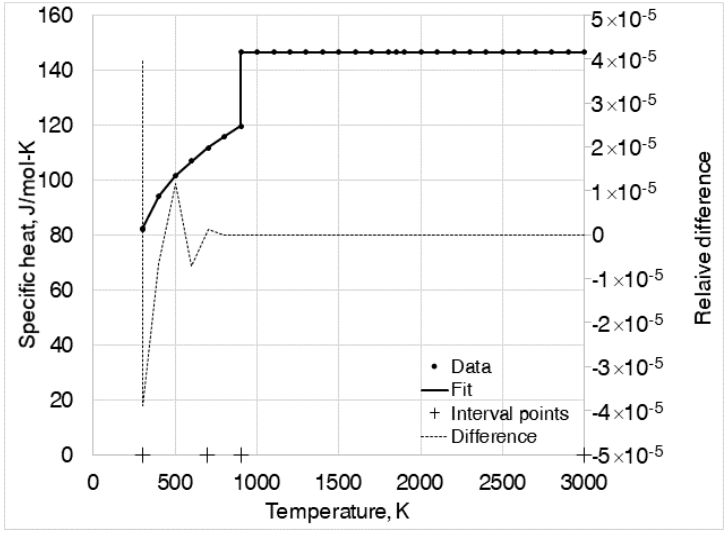


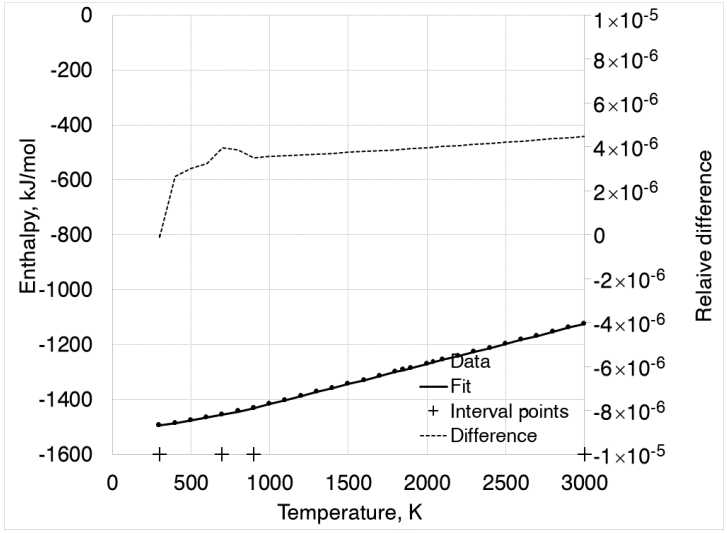


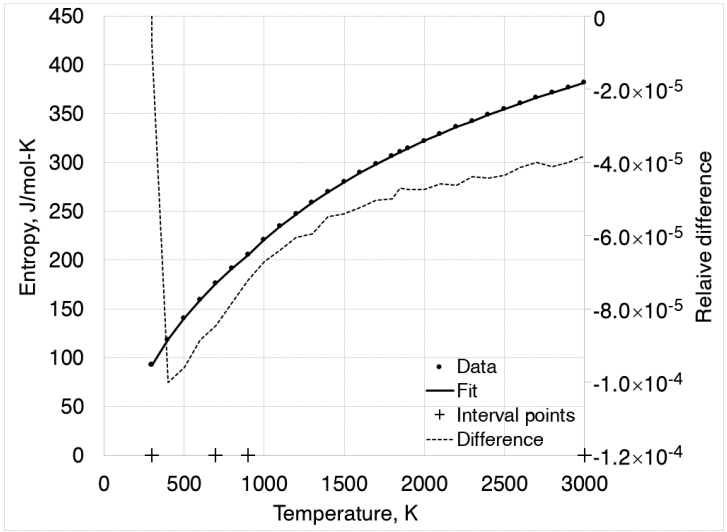


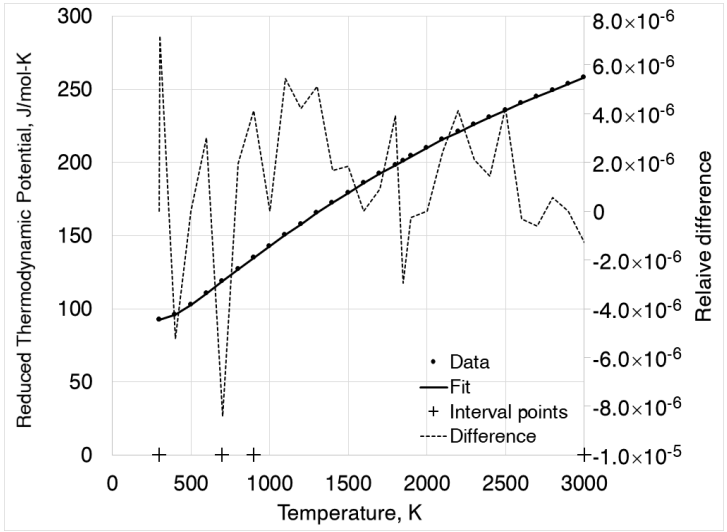
Magnesium Silicate ($MgSiO_3$) – Mg1O3Si1(l)

MW = 100.389 g/mol – $\Delta_f H^o$(298.15 K) = -1494.862 kJ/mol – S^o(298.15 K) = 92.515 J/mol-K – p^o = 0.1 MPa



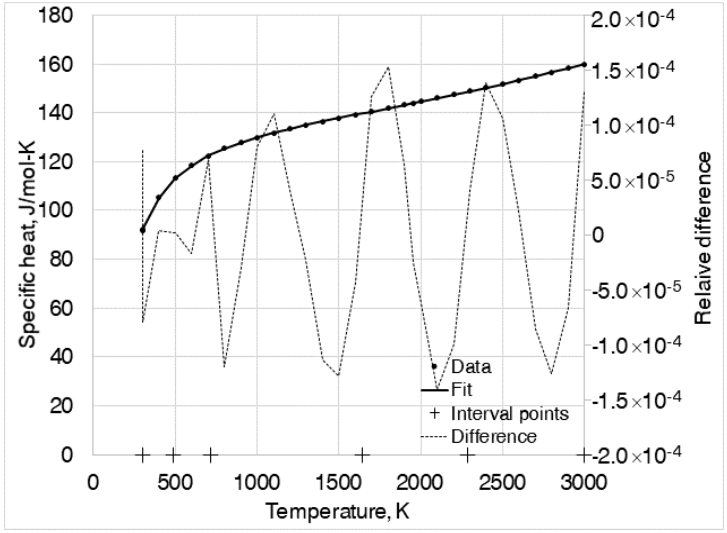


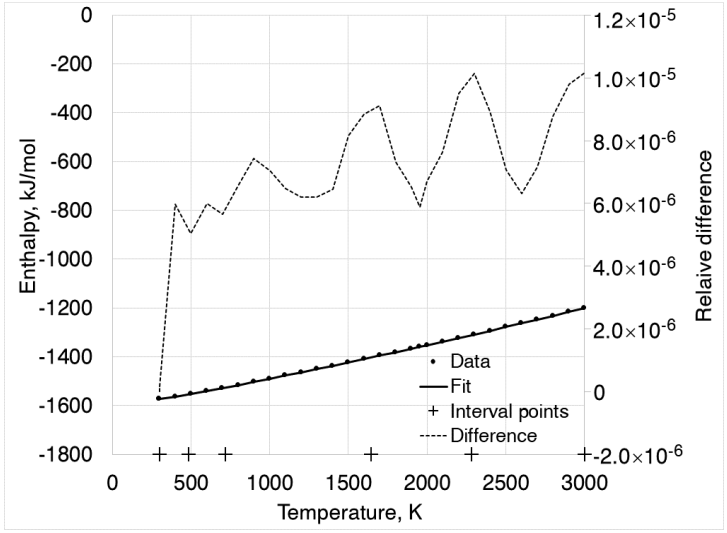


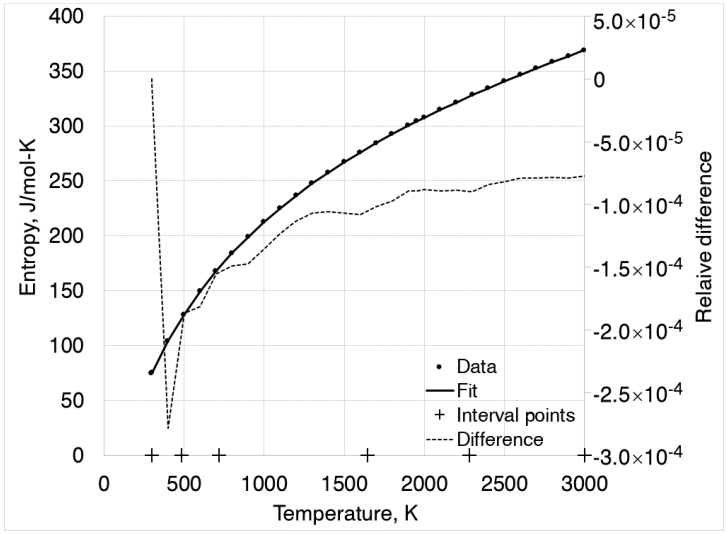


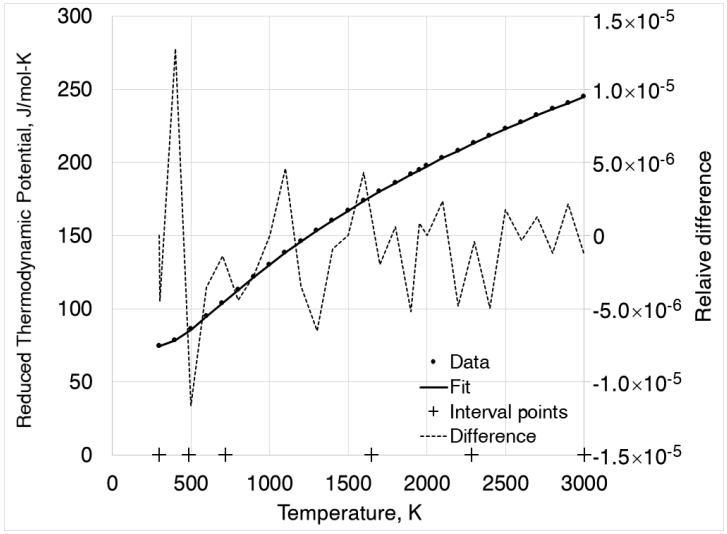
Magnesium Titanium Oxide ($MgTiO_3$) – Mg1O3Ti1(cr)

MW = 120.18 g/mol – $\Delta_f H^o$(298.15 K) = -1572.556 kJ/mol – S^o(298.15 K) = 74.559 J/mol-K – p^o = 0.1 MPa



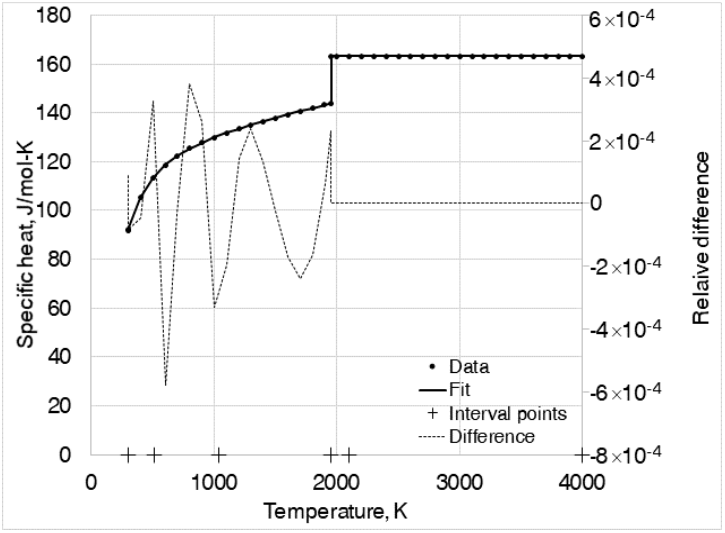


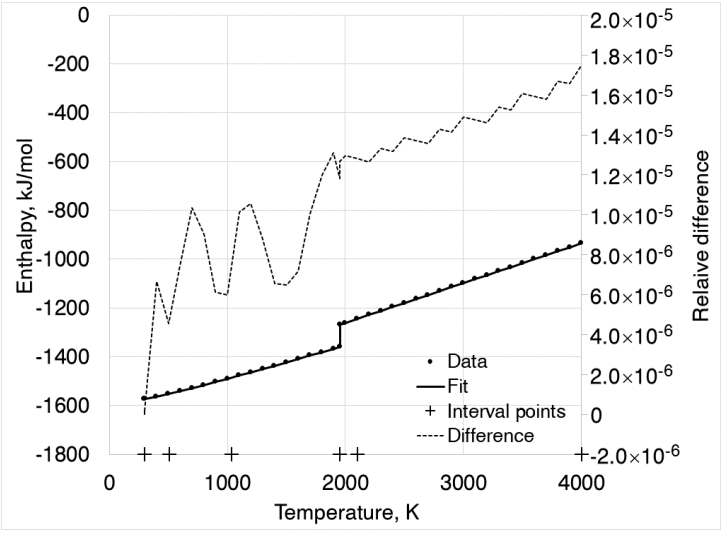


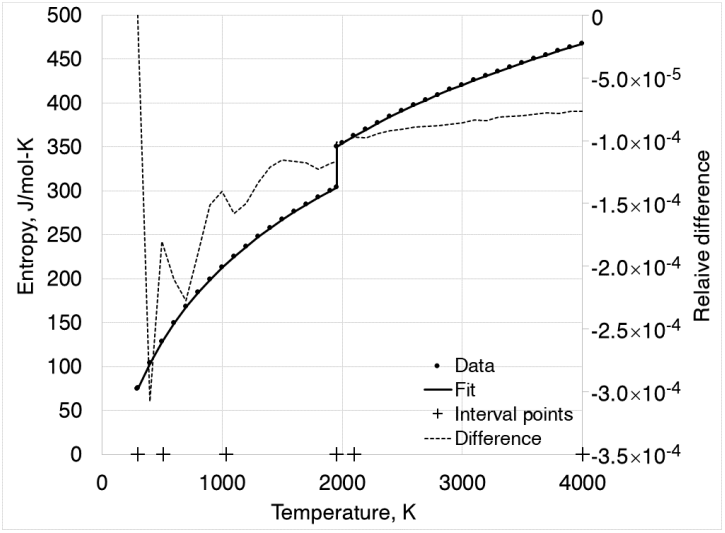


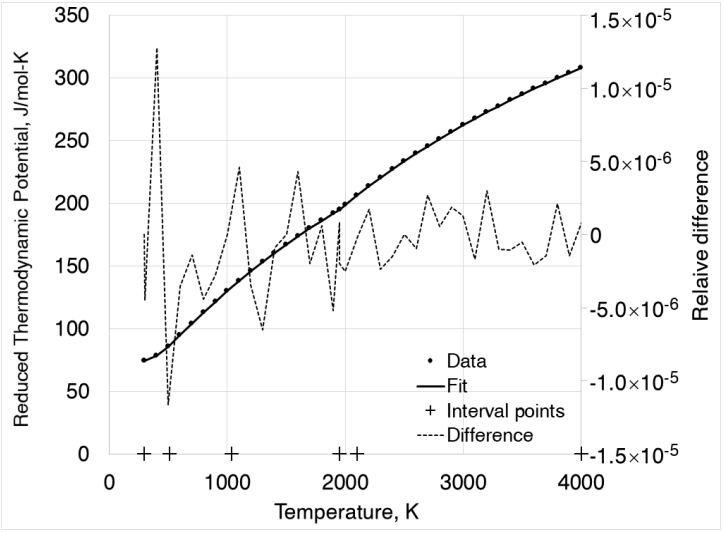
Magnesium Titanium Oxide ($MgTiO_3$) – Mg1O3Ti1(cr,l)

MW = 120.18 g/mol – $\Delta_f H^o$(298.15 K) = -1572.556 kJ/mol – S^o(298.15 K) = 74.559 J/mol-K – p^o = 0.1 MPa



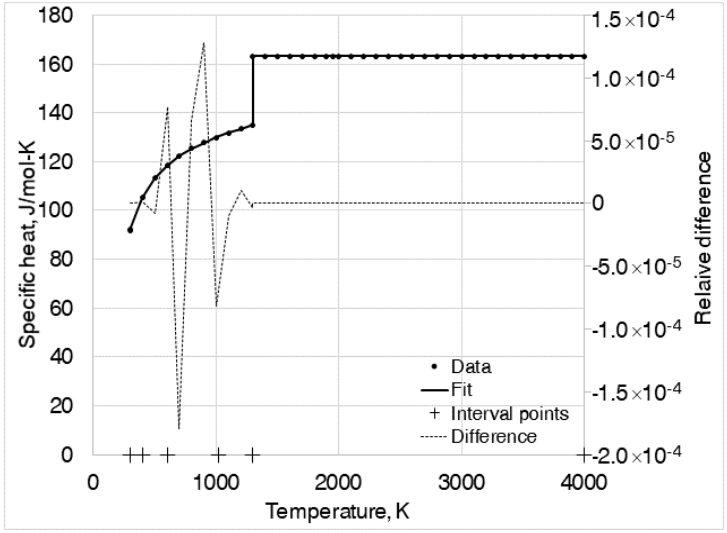


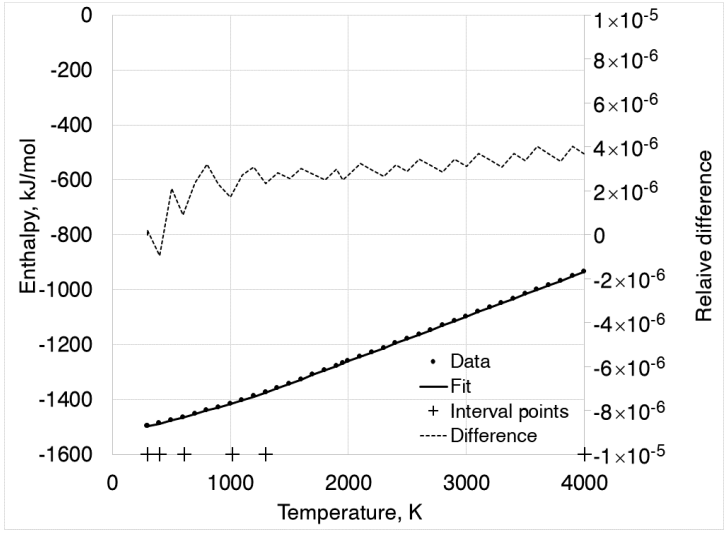


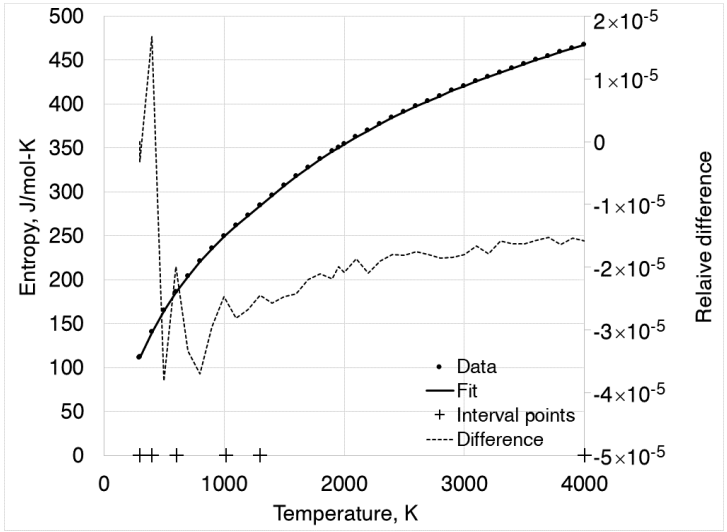


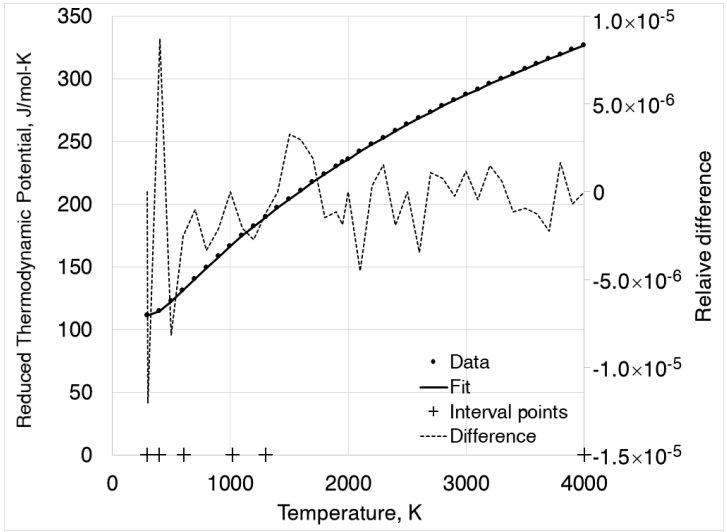
Magnesium Titanium Oxide ($MgTiO_3$) – Mg1O3Ti1(l)

MW = 120.18 g/mol – $\Delta_f H^o$(298.15 K) = -1497.631 kJ/mol – S^o(298.15 K) = 111.082 J/mol-K – p^o = 0.1 MPa



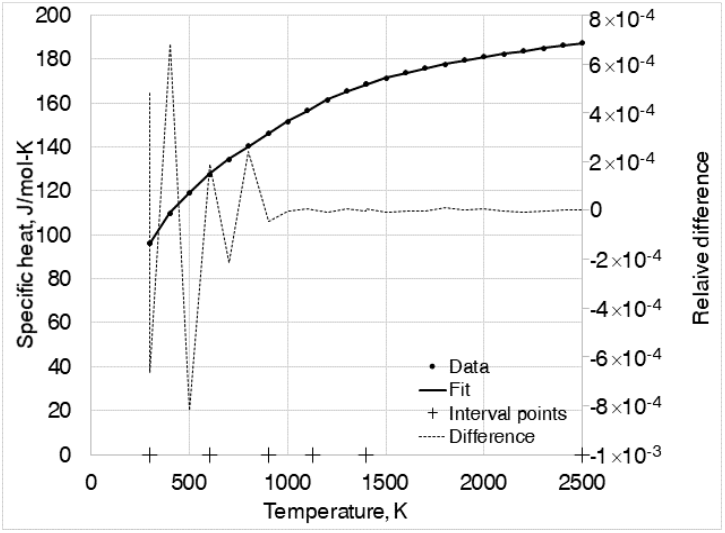


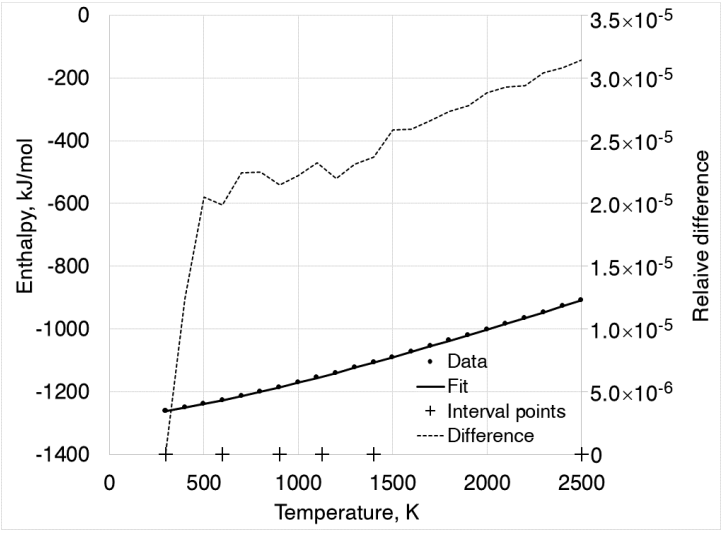


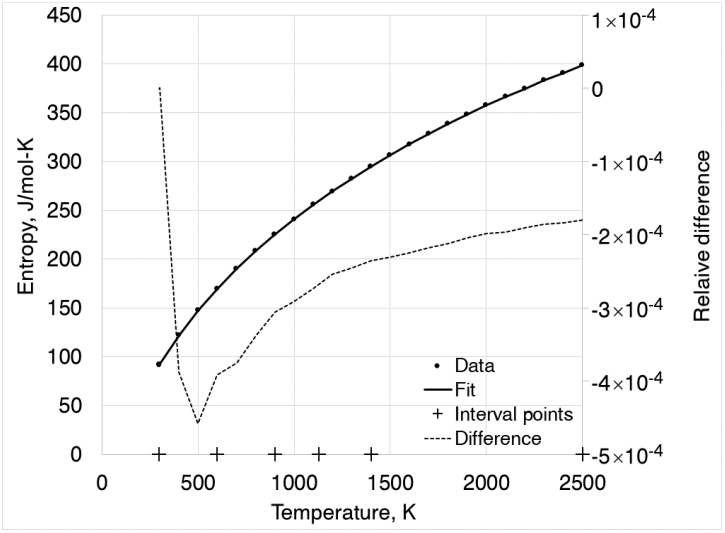


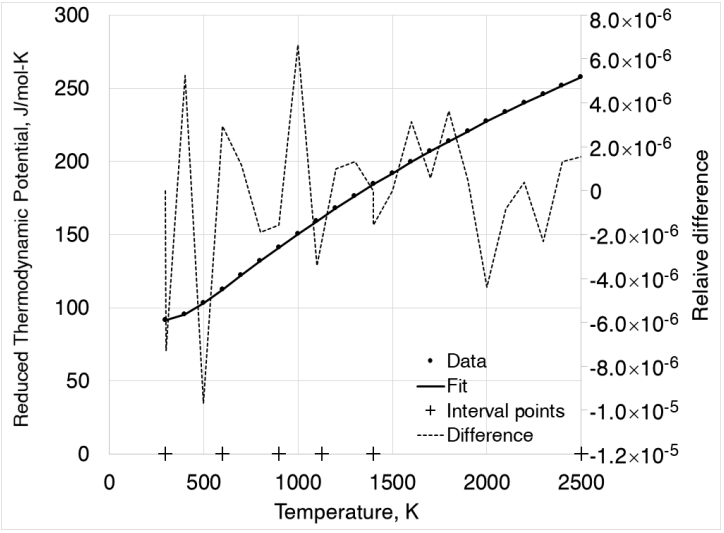
Magnesium Sulfate ($MgSO_4$) – Mg1O4S1(cr)

MW = 120.36 g/mol – $\Delta_f H^o$(298.15 K) = -1261.786 kJ/mol – S^o(298.15 K) = 91.396 J/mol-K – p^o = 0.1 MPa



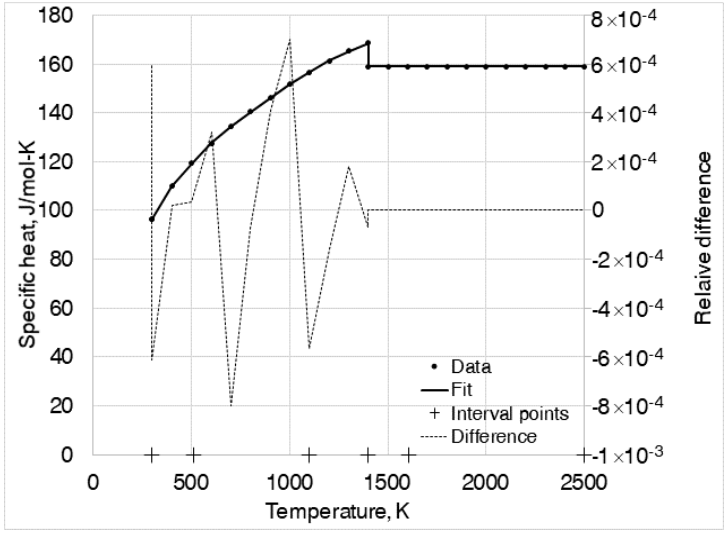


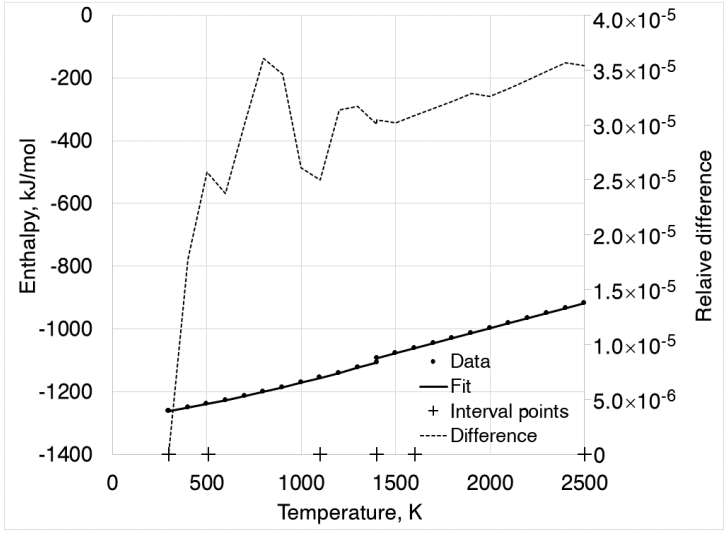


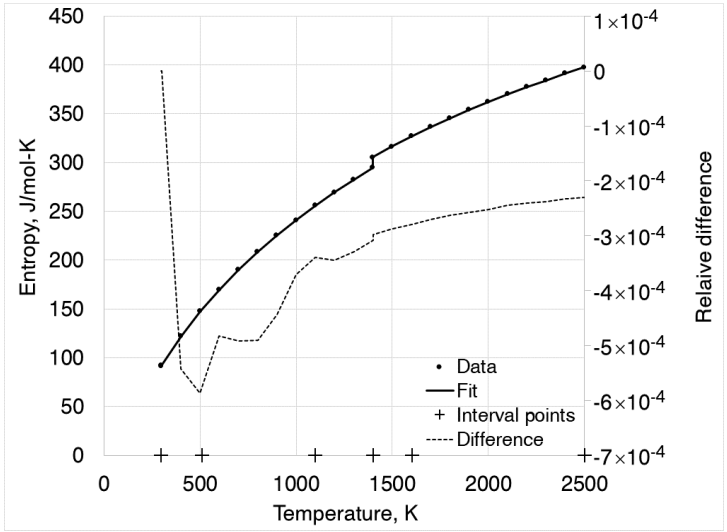


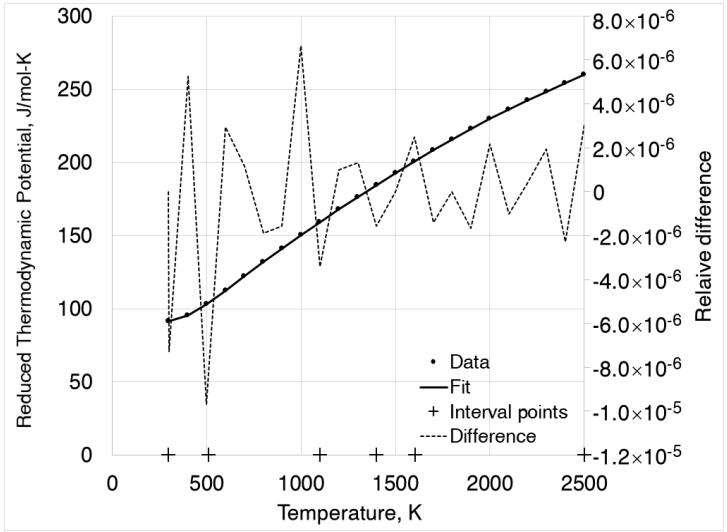
Magnesium Sulfate ($MgSO_4$) – Mg1O4S1(cr,l)

MW = 120.36 g/mol – $\Delta_f H^o$(298.15 K) = -1261.786 kJ/mol – S^o(298.15 K) = 91.396 J/mol-K – p^o = 0.1 MPa



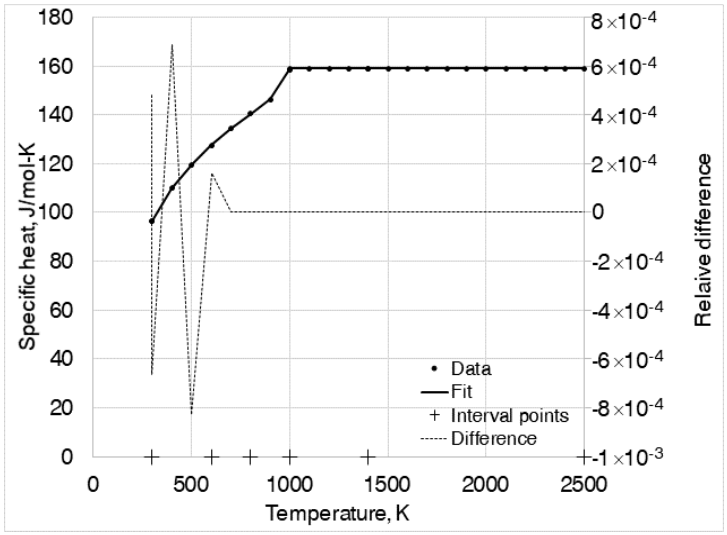


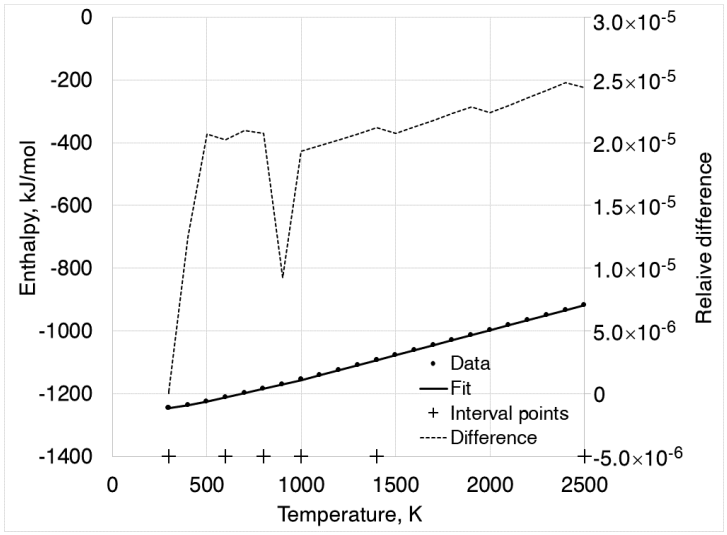


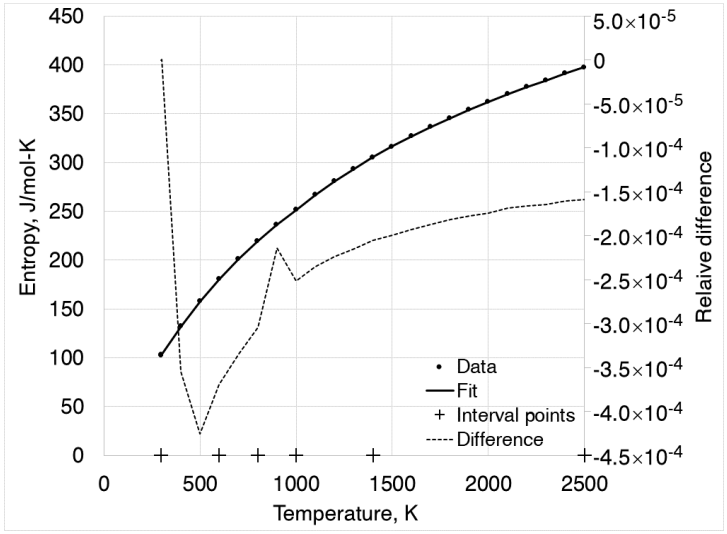


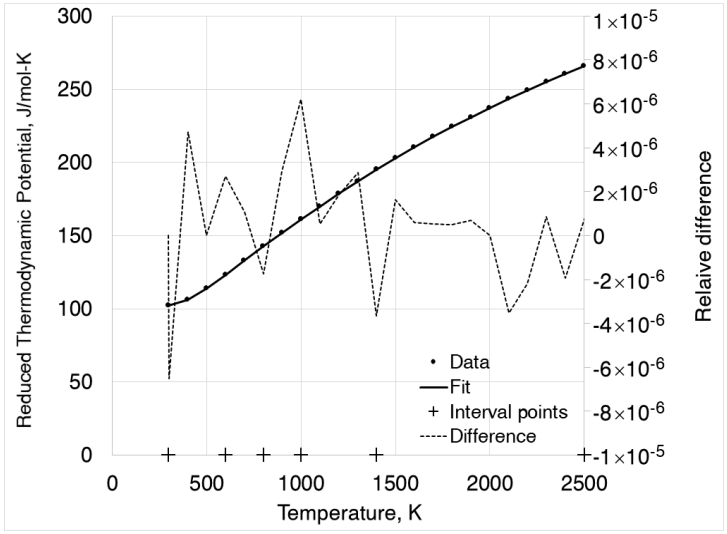
Magnesium Sulfate ($MgSO_4$) – Mg1O4S1(l)

MW = 120.36 g/mol – $\Delta_f H^o$(298.15 K) = -1246.587 kJ/mol – S^o(298.15 K) = 102.125 J/mol-K – p^o = 0.1 MPa



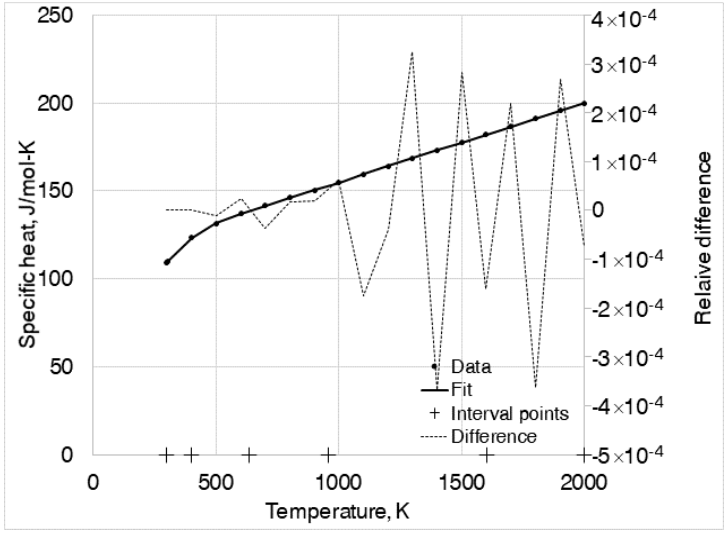


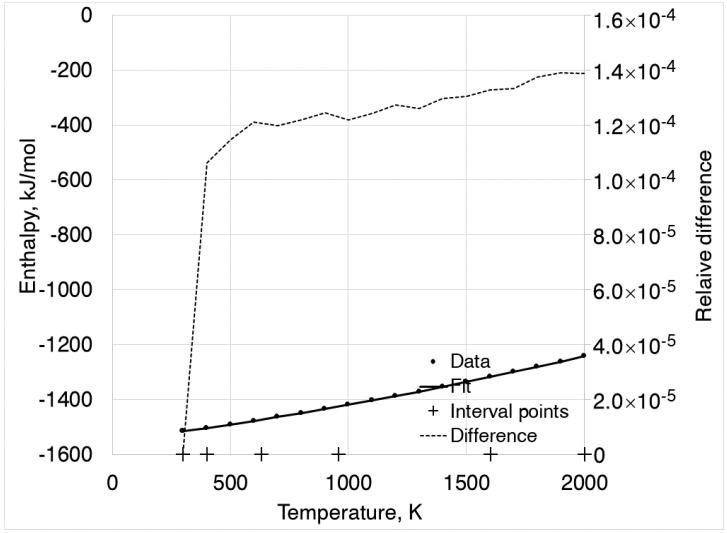


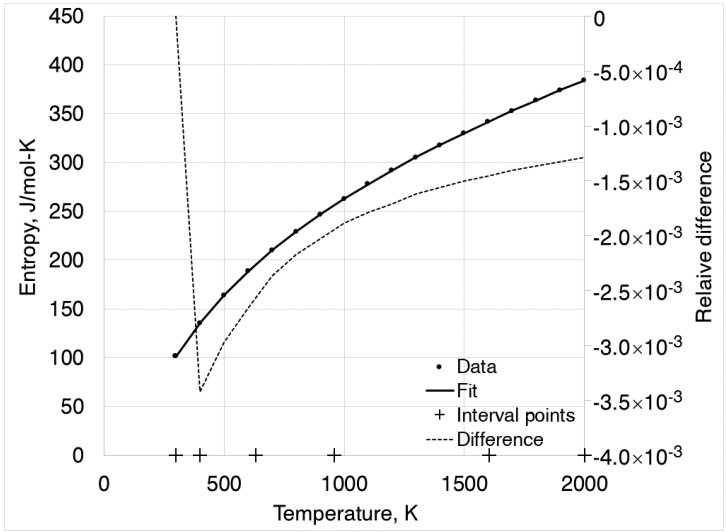


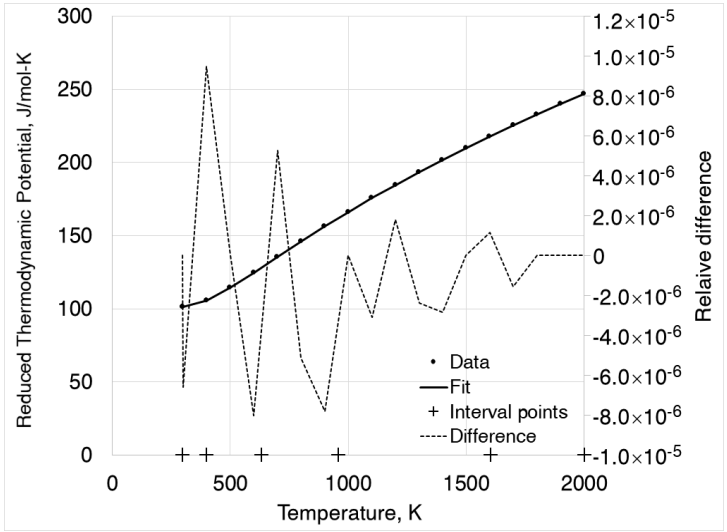
Magnesium Tungsten Oxide ($MgWO_4$) – Mg1O4W1(cr)

MW = 272.15 g/mol – $\Delta_f H^o$(298.15 K) = -1515.863 kJ/mol – S^o(298.15 K) = 101.186 J/mol-K – p^o = 0.1 MPa



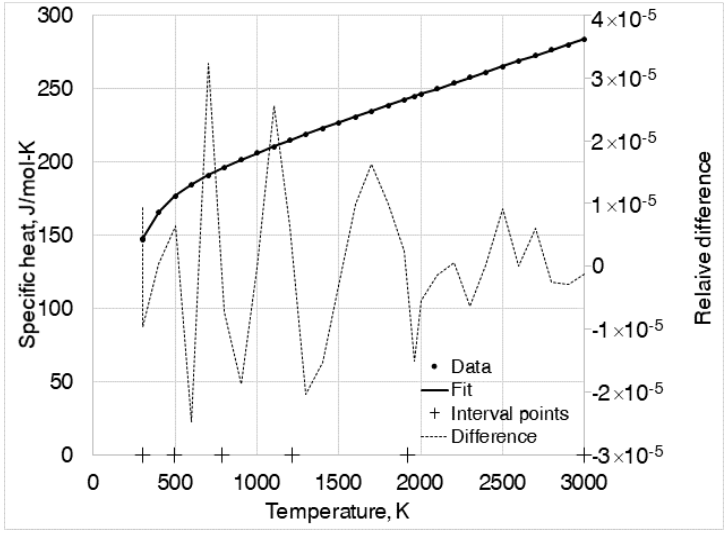


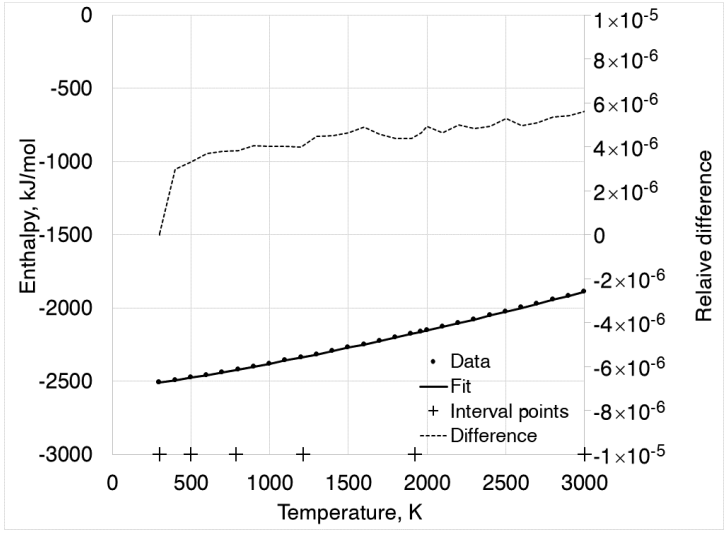


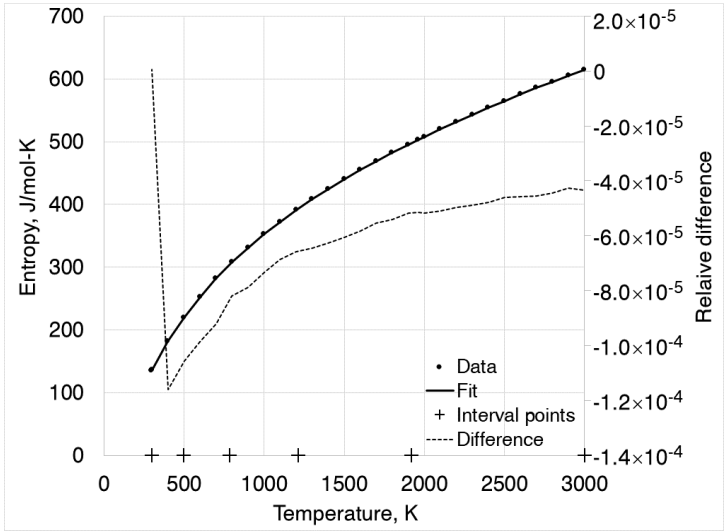


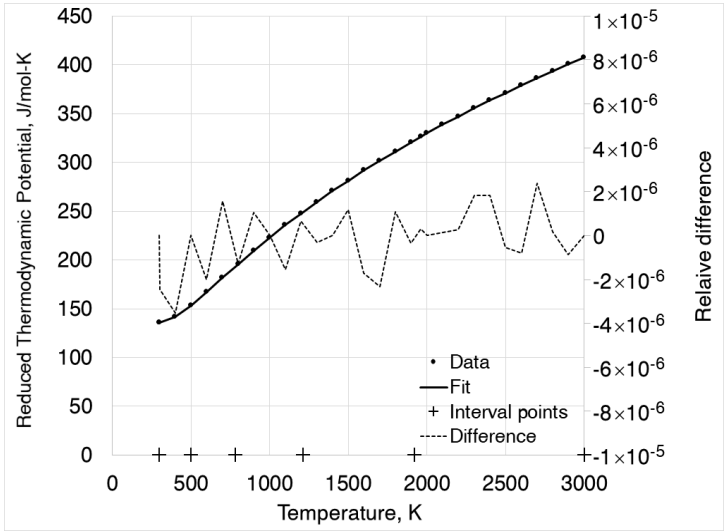
Magnesium Titanium Oxide ($MgTi_2O_5$) – Mg1O5Ti2(cr)

MW = 200.06 g/mol – $\Delta_f H^o$(298.15 K) = -2509.354 kJ/mol – S^o(298.15 K) = 135.603 J/mol-K – p^o = 0.1 MPa



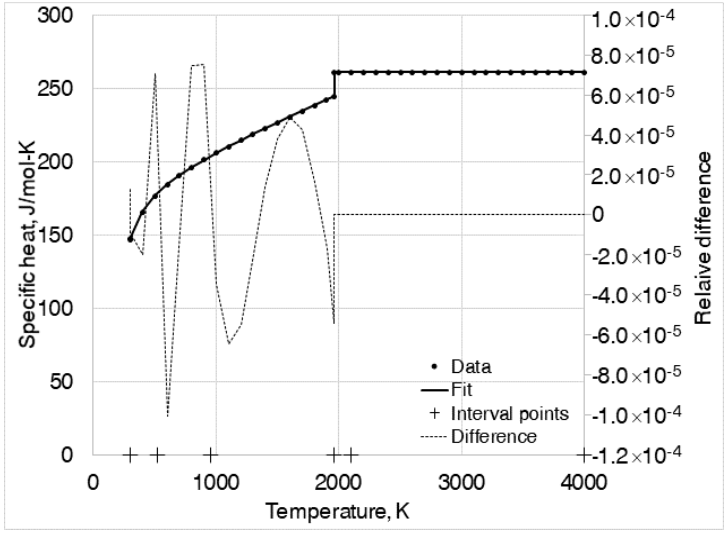


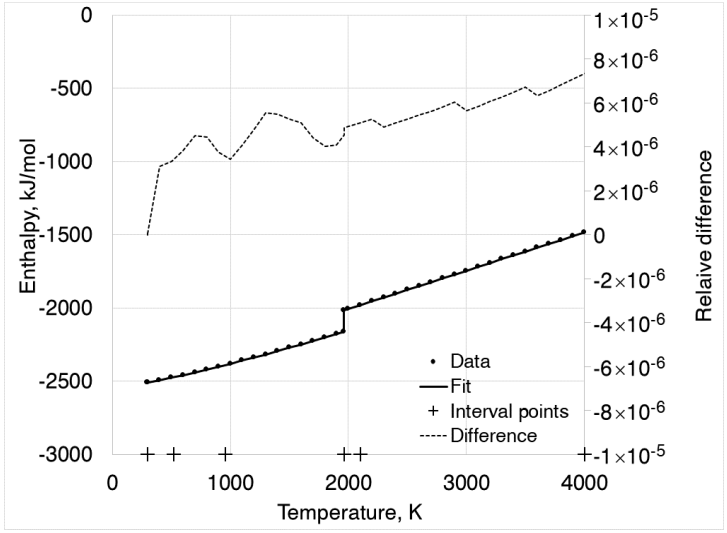


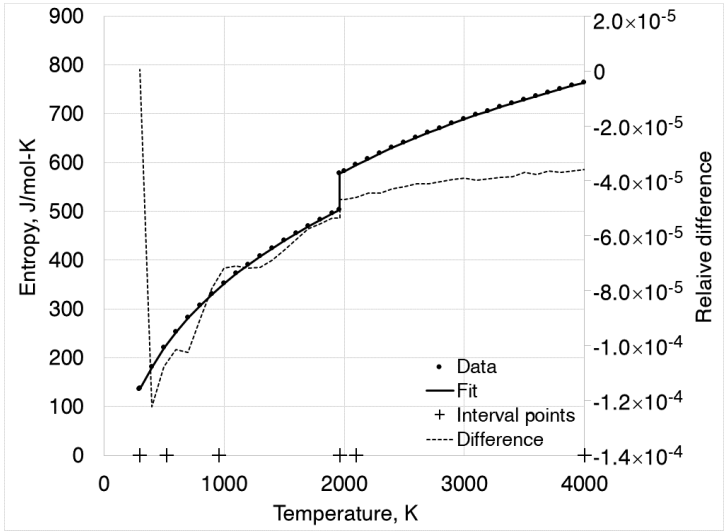


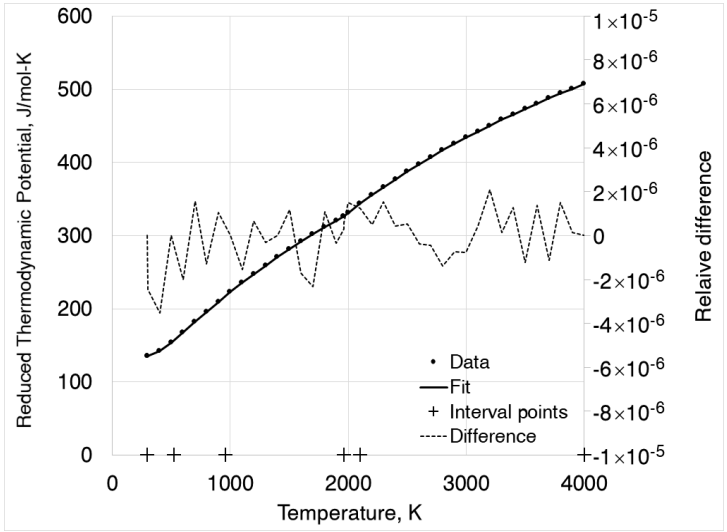
Magnesium Titanium Oxide ($MgTi_2O_5$) – Mg1O5Ti2(cr,l)

MW = 200.06 g/mol – $\Delta_f H^o$(298.15 K) = -2509.354 kJ/mol – S^o(298.15 K) = 135.603 J/mol-K – p^o = 0.1 MPa



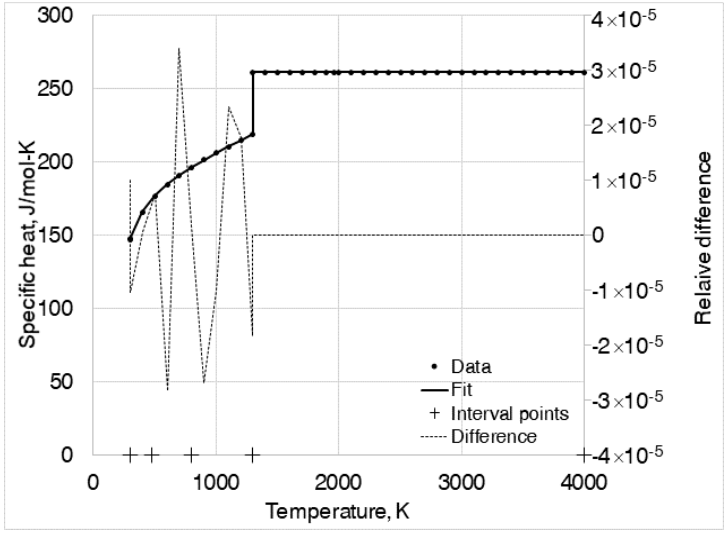


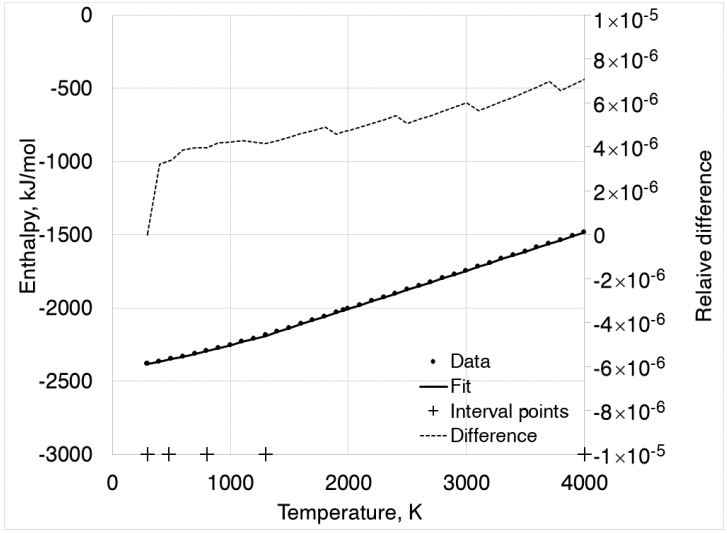


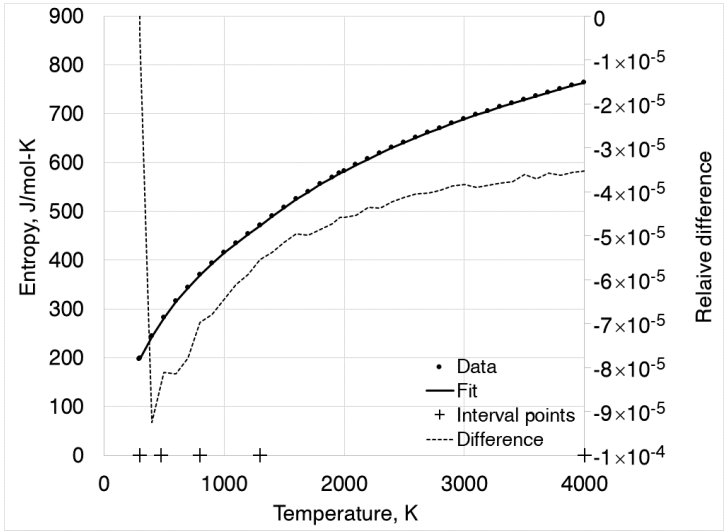


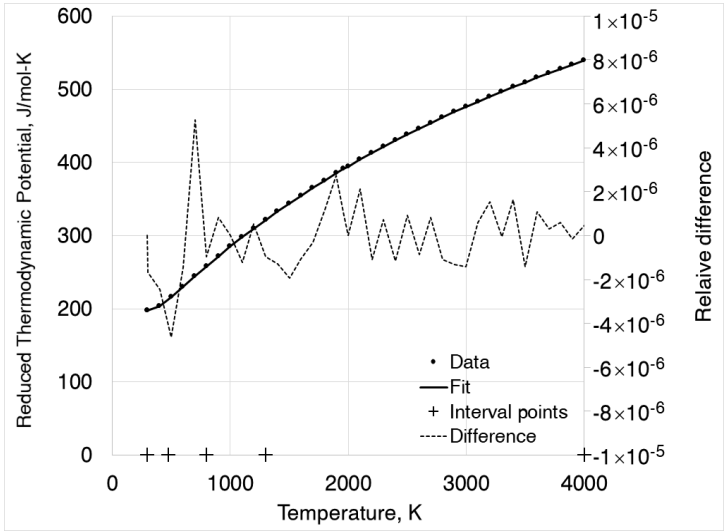
Magnesium Titanium Oxide ($MgTi_2O_5$) – Mg1O5Ti2(l)

MW = 200.06 g/mol – $\Delta_f H^o$(298.15 K) = -2382.307 kJ/mol – S^o(298.15 K) = 197.783 J/mol-K – p^o = 0.1 MPa



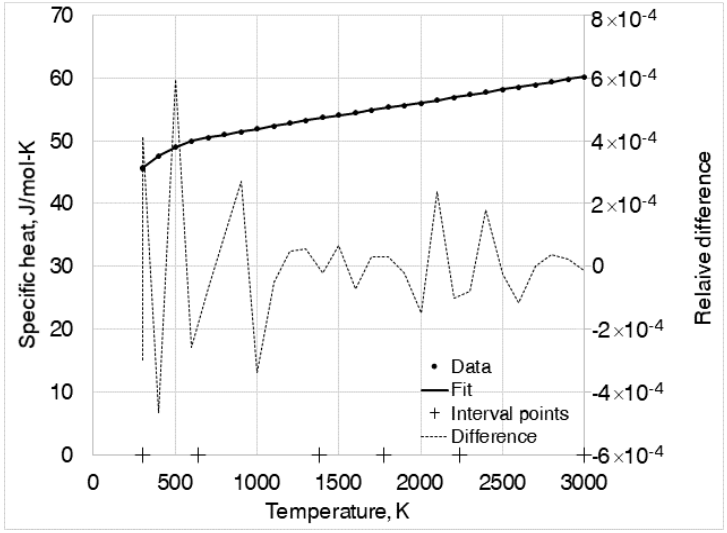


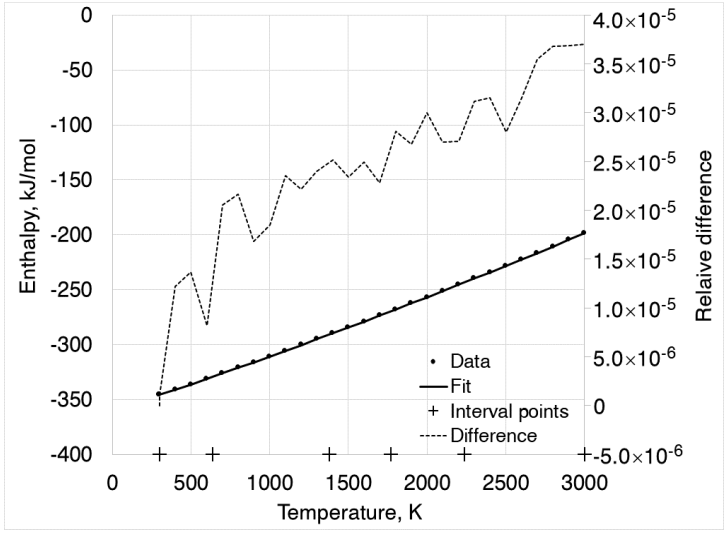


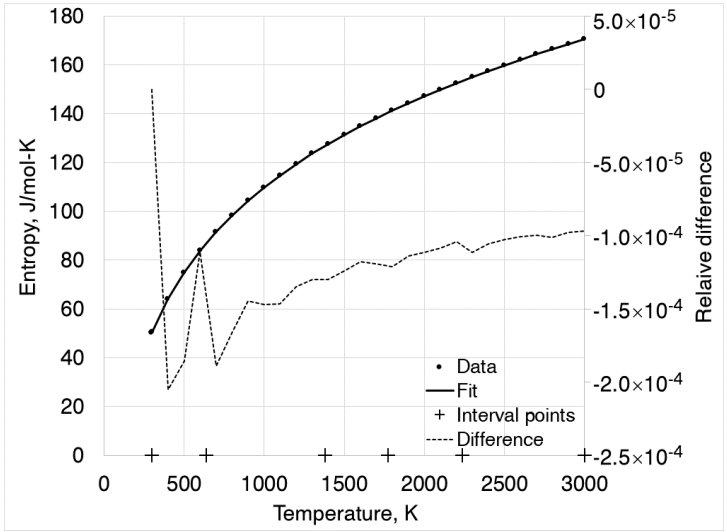


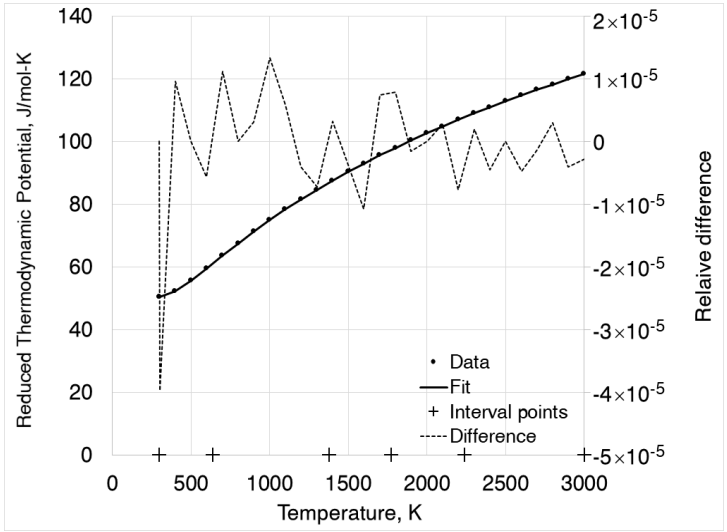
Magnesium Sulfide (MgS) – Mg1S1(cr)

MW = 56.36 g/mol – $\Delta_f H^o$(298.15 K) = -345.724 kJ/mol – S^o(298.15 K) = 50.334 J/mol-K – p^o = 0.1 MPa



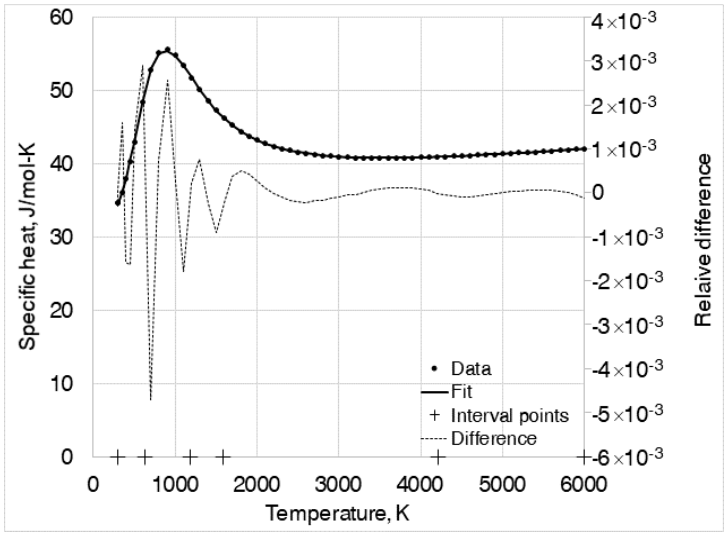


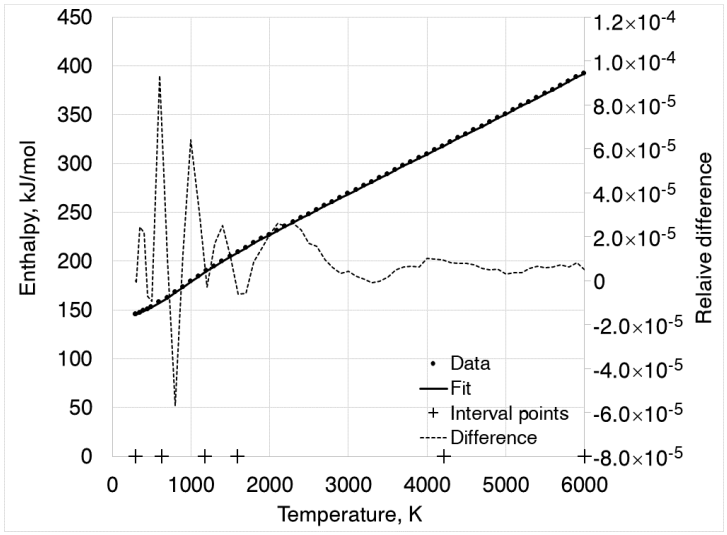


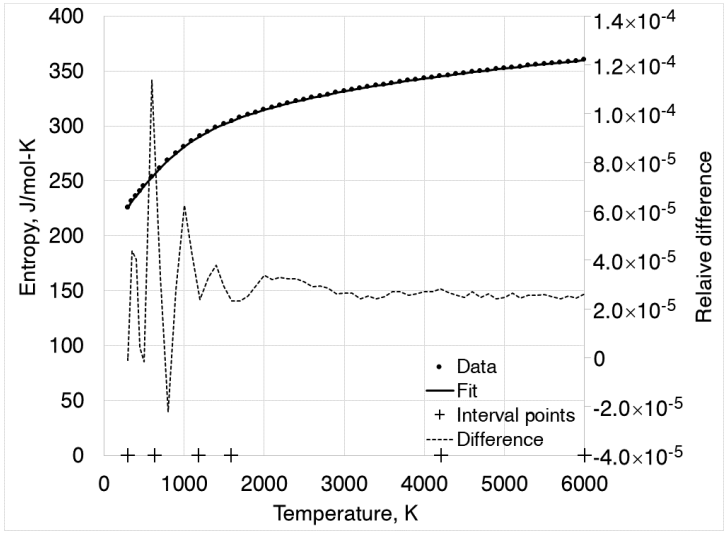


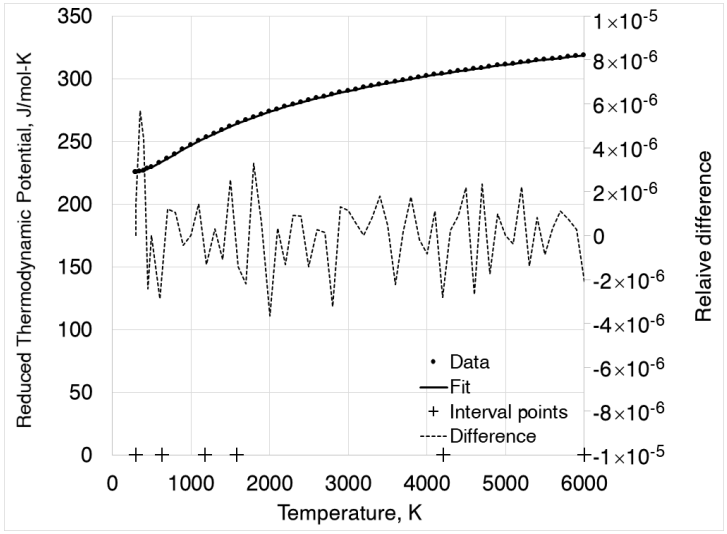
Magnesium Sulfide (MgS) – Mg1S1(g)

MW = 56.36 g/mol – $\Delta_f H^o$(298.15 K) = 145.227 kJ/mol – S^o(298.15 K) = 225.483 J/mol-K – p^o = 0.1 MPa



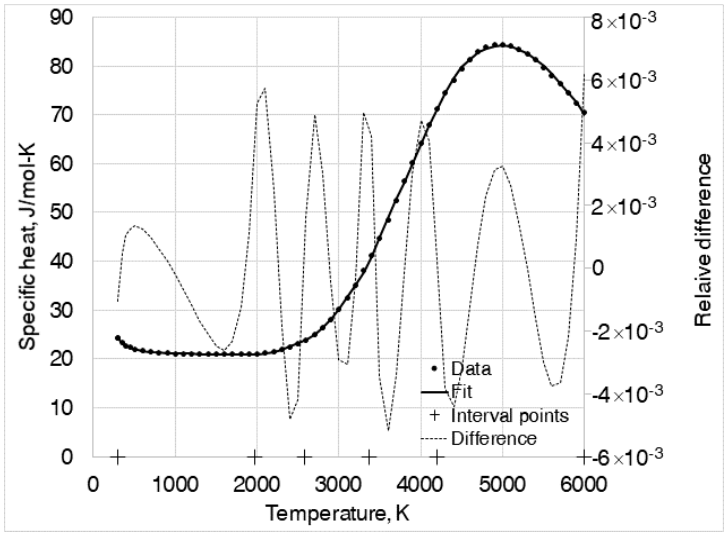


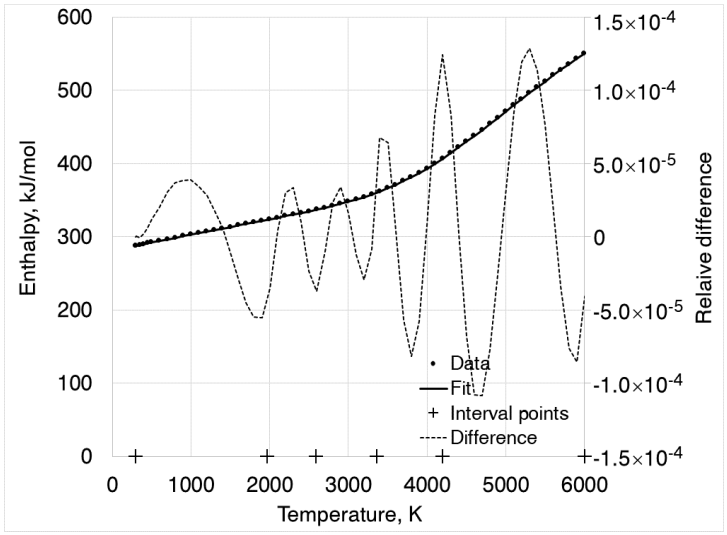


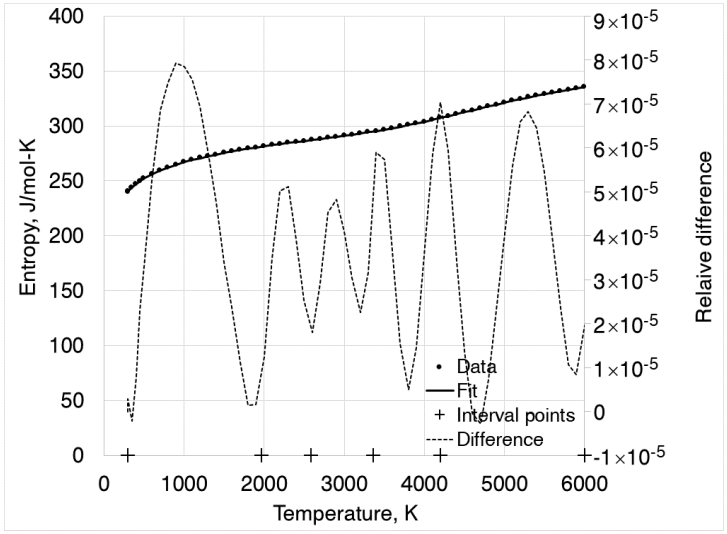


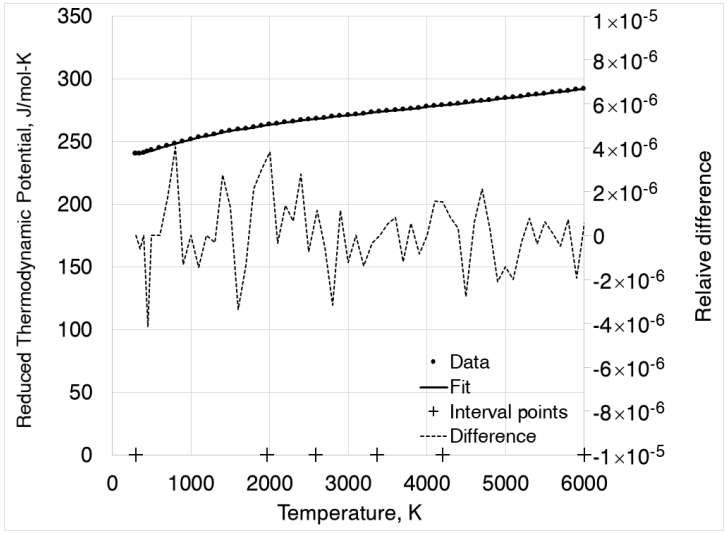
Magnesium (Mg_2) – Mg2(g)

MW = 48.61 g/mol – $\Delta_f H^o$(298.15 K) = 287.625 kJ/mol – S^o(298.15 K) = 240.186 J/mol-K – p^o = 0.1 MPa



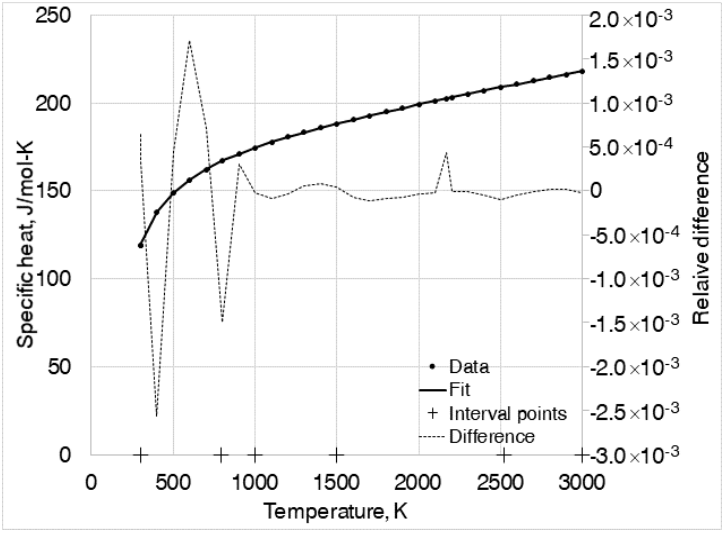


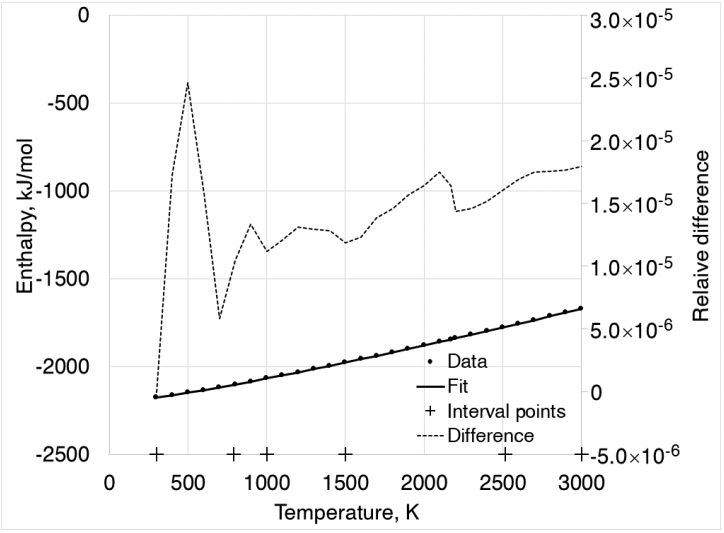


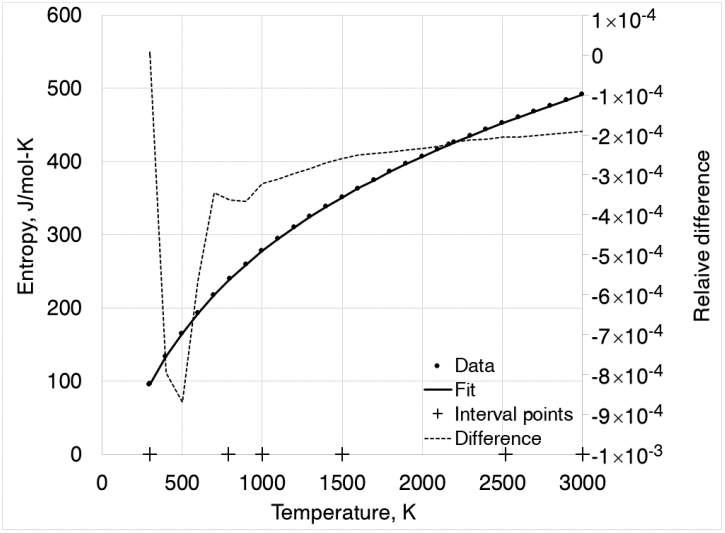


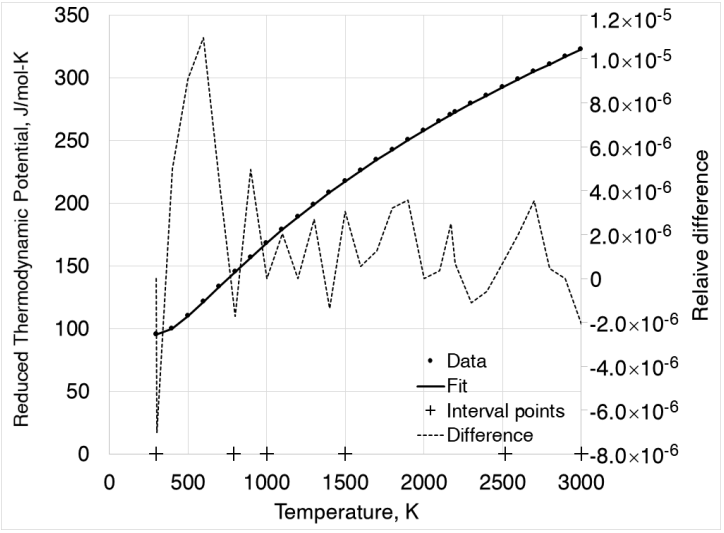
Magnesium Silicate (Mg_2SiO_4) – Mg2O4Si1(cr)

MW = 140.693 g/mol – $\Delta_f H^o$(298.15 K) = -2176.935 kJ/mol – S^o(298.15 K) = 95.14 J/mol-K – p^o = 0.1 MPa



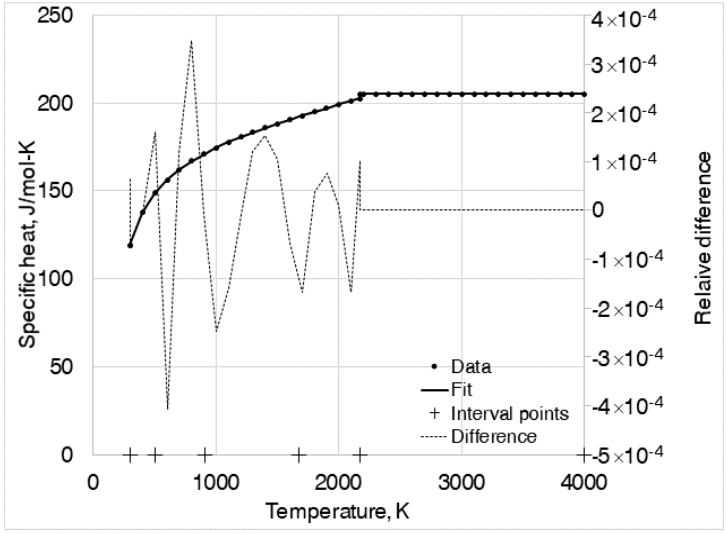


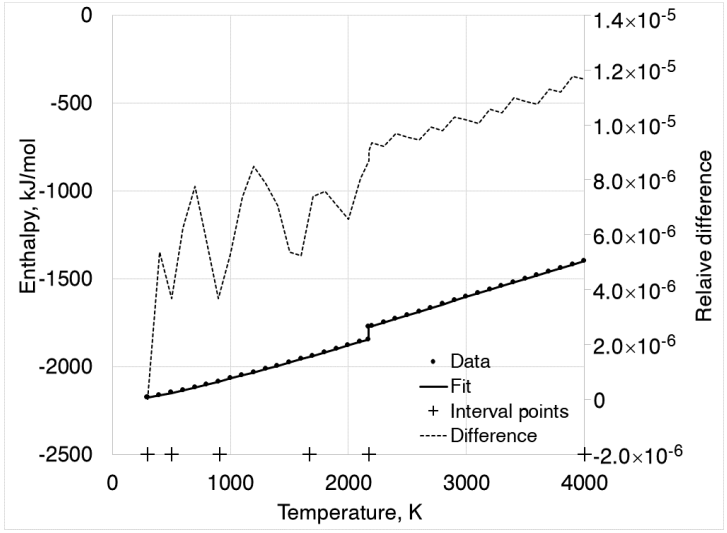


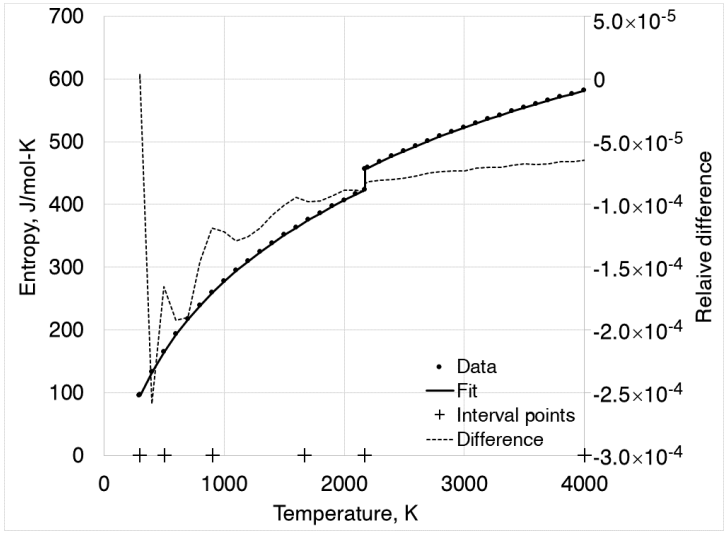


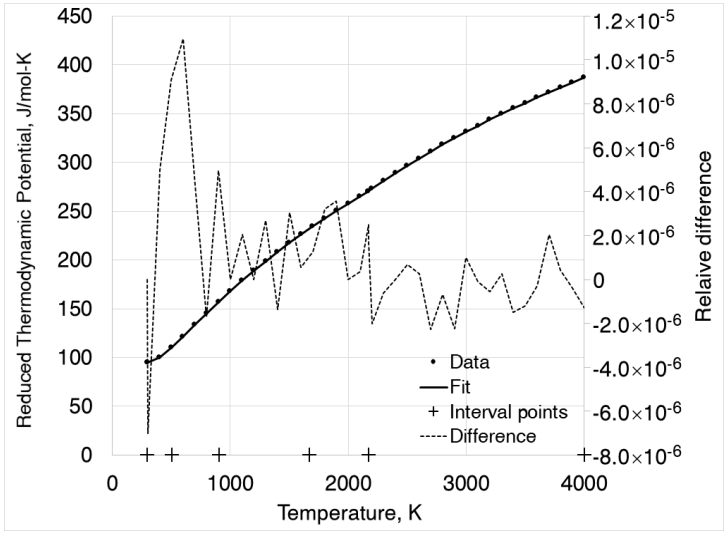
Magnesium Silicate (Mg_2SiO_4) – Mg2O4Si1(cr,l)

MW = 140.693 g/mol – $\Delta_f H^o$(298.15 K) = -2176.935 kJ/mol – S^o(298.15 K) = 95.14 J/mol-K – p^o = 0.1 MPa



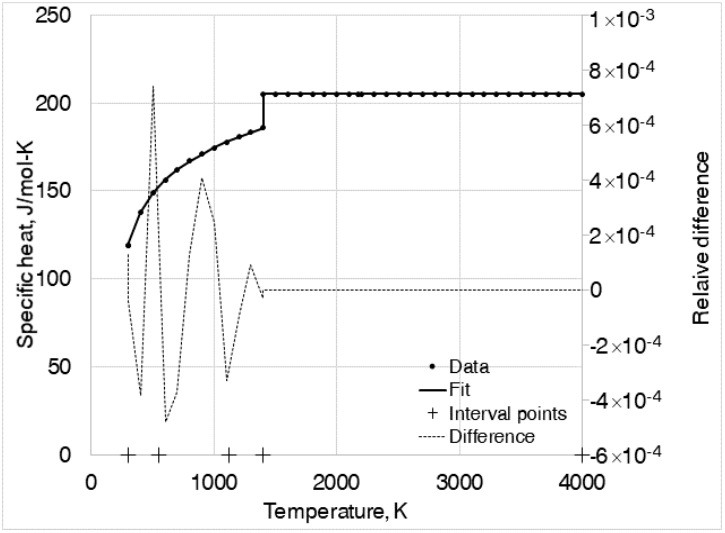


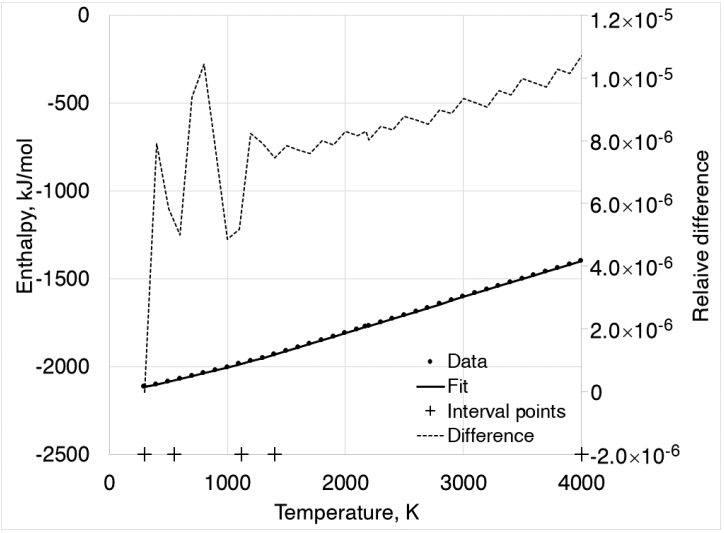


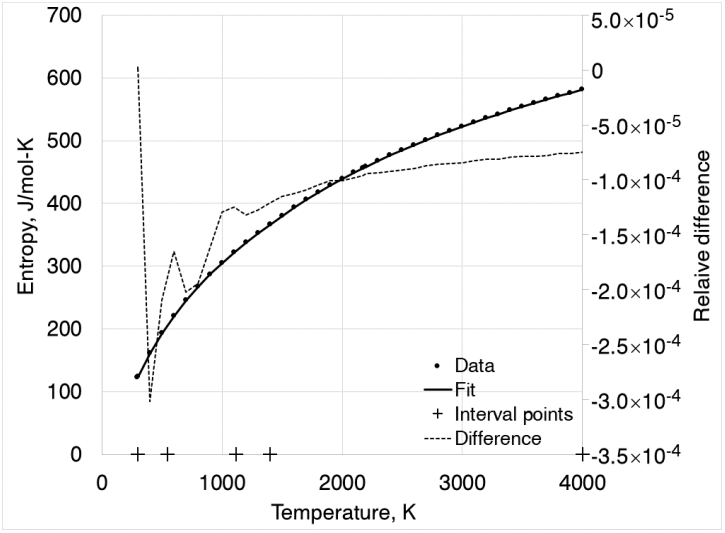


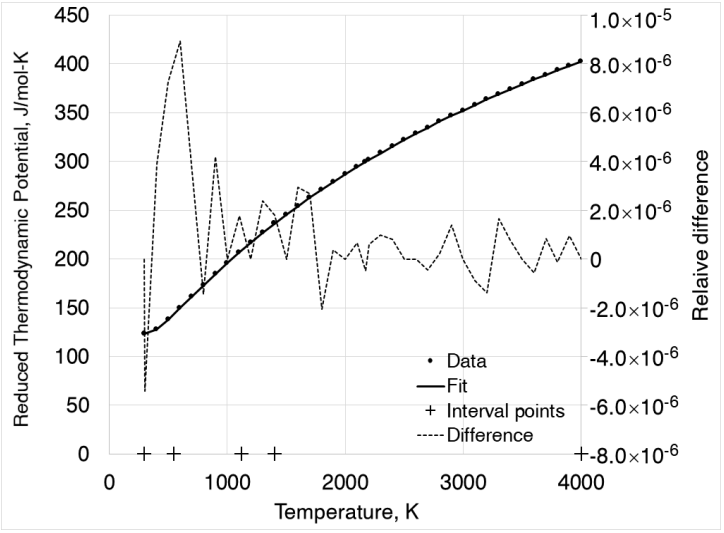
Magnesium Silicate (Mg_2SiO_4) – Mg2O4Si1(l)

MW = 140.693 g/mol – $\Delta_f H^o$(298.15 K) = -2113.878 kJ/mol – S^o(298.15 K) = 123.044 J/mol-K – p^o = 0.1 MPa



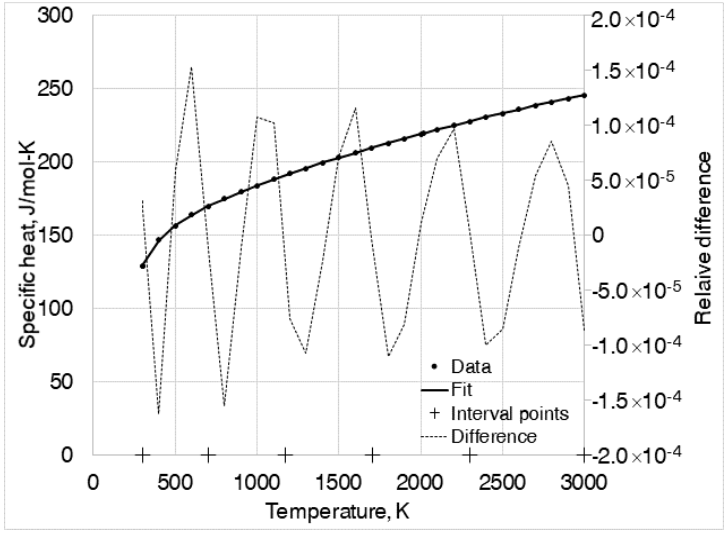


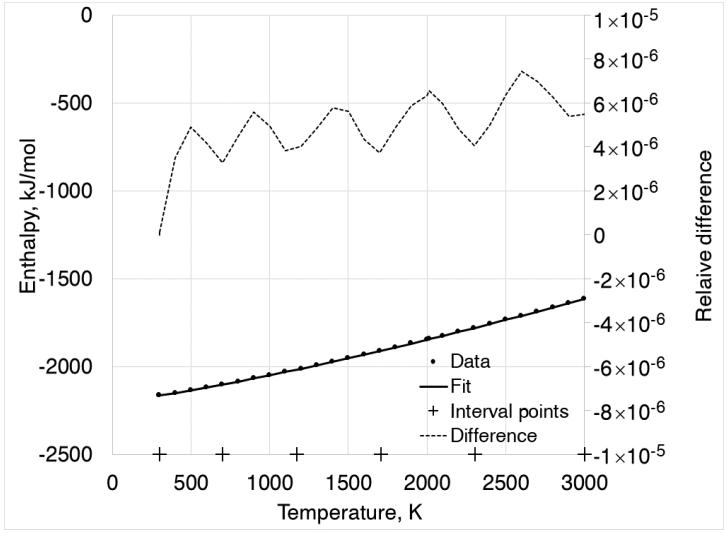


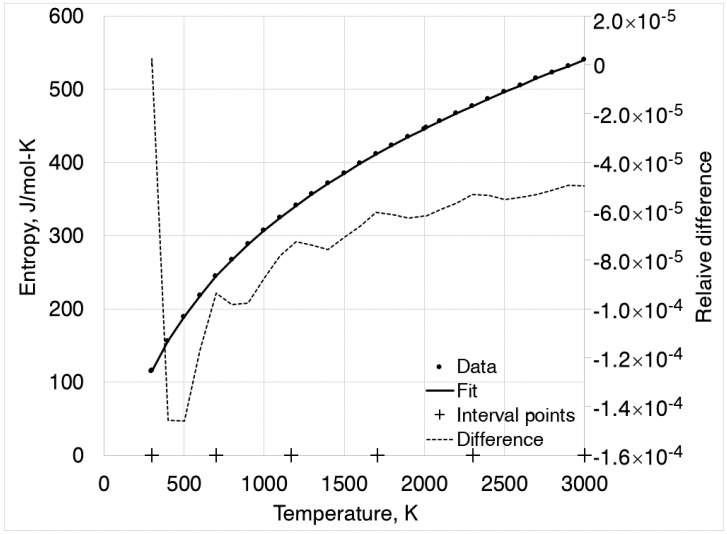


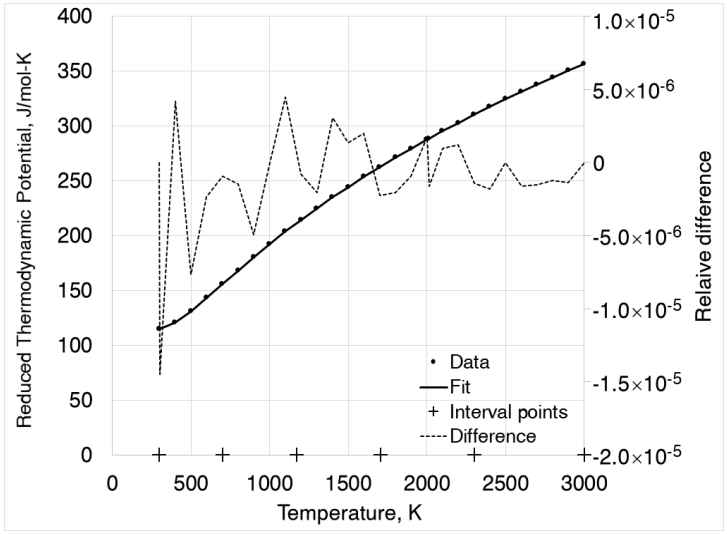
Magnesium Titanium Oxide (Mg_2TiO_4) – Mg2O4Ti1(cr)

MW = 160.49 g/mol – $\Delta_f H^o$(298.15 K) = -2164.383 kJ/mol – S^o(298.15 K) = 115.102 J/mol-K – p^o = 0.1 MPa



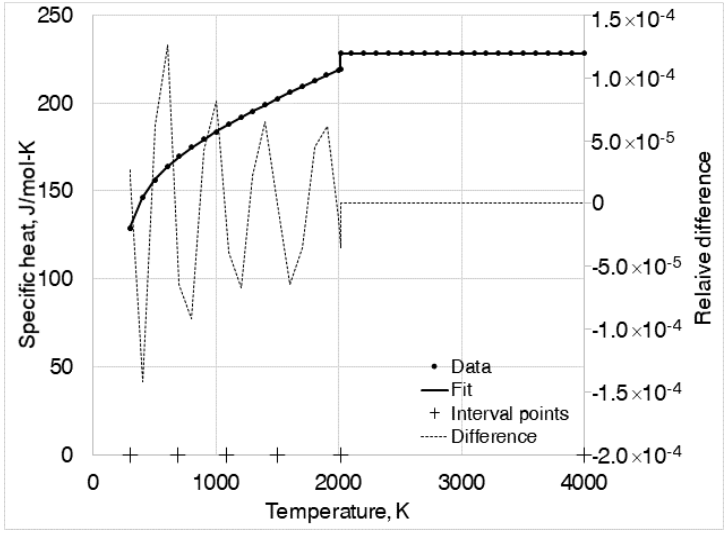


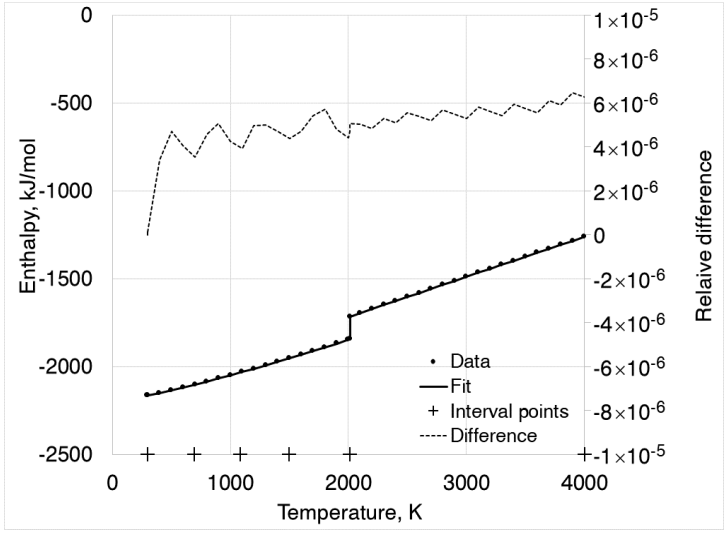


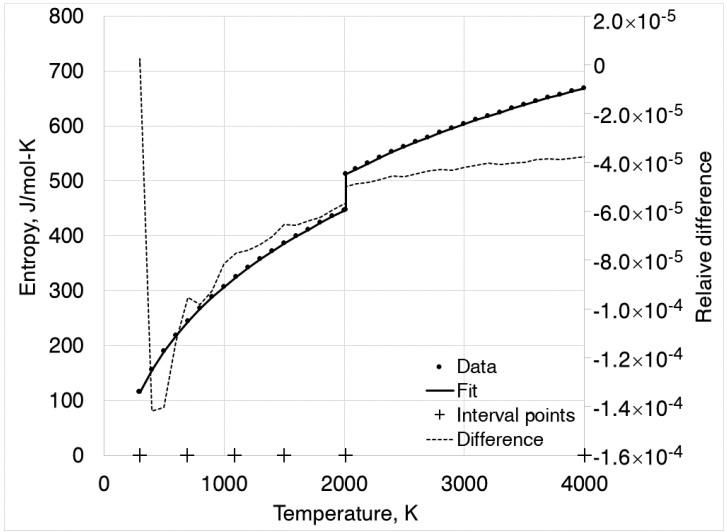


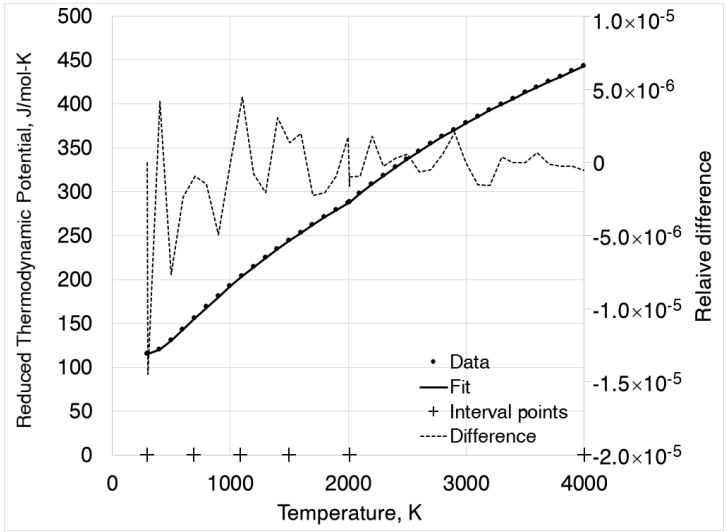
Magnesium Titanium Oxide (Mg_2TiO_4) – Mg2O4Ti1(cr,l)

MW = 160.49 g/mol – $\Delta_f H^o$(298.15 K) = -2164.383 kJ/mol – S^o(298.15 K) = 115.102 J/mol-K – p^o = 0.1 MPa



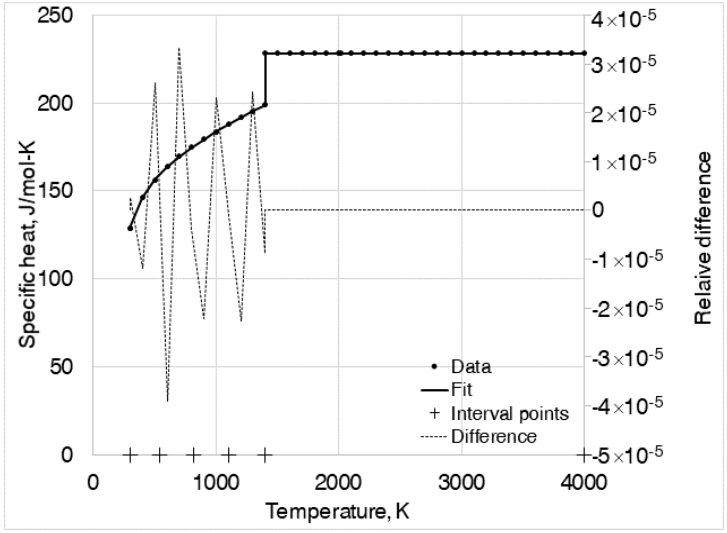


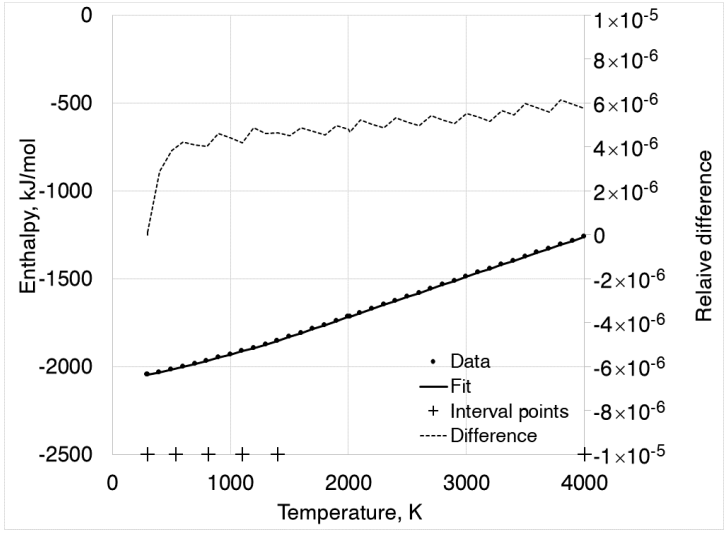


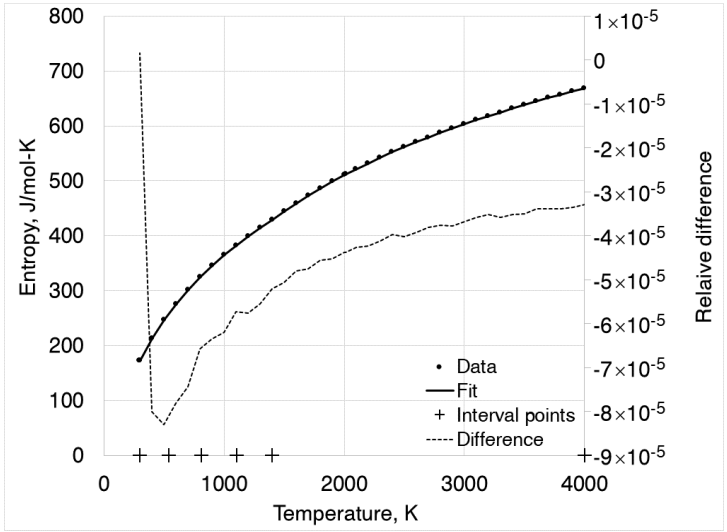


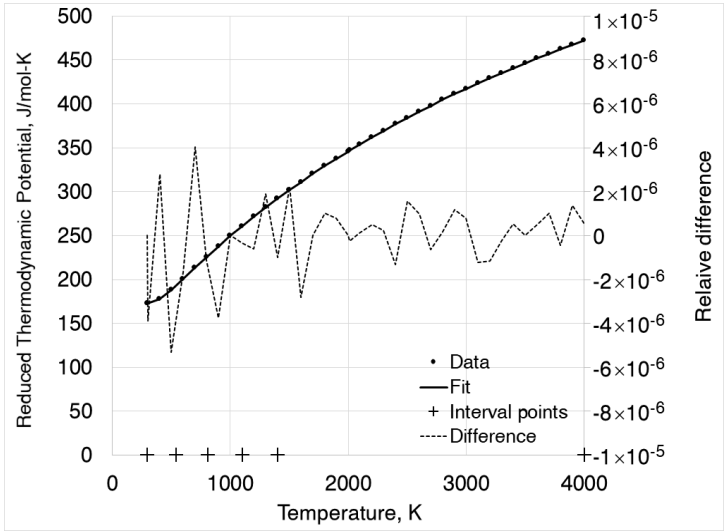
Magnesium Titanium Oxide (Mg_2TiO_4) – Mg2O4Ti1(l)

MW = 160.49 g/mol – $\Delta_f H^o$(298.15 K) = -2046.329 kJ/mol – S^o(298.15 K) = 172.412 J/mol-K – p^o = 0.1 MPa



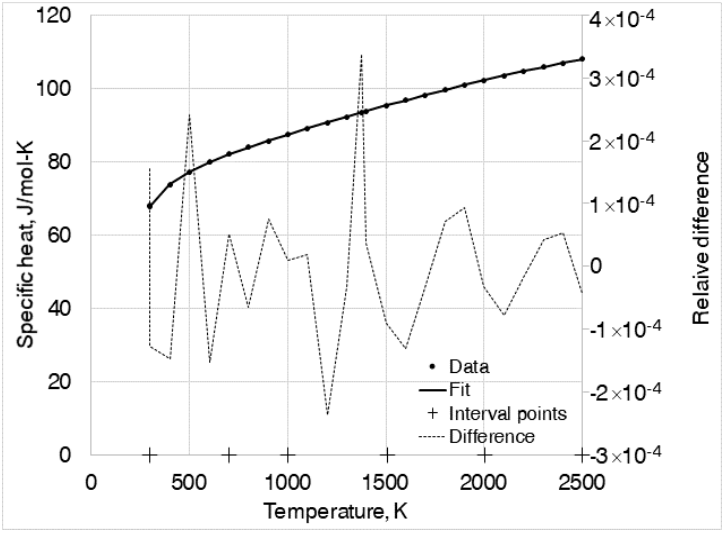


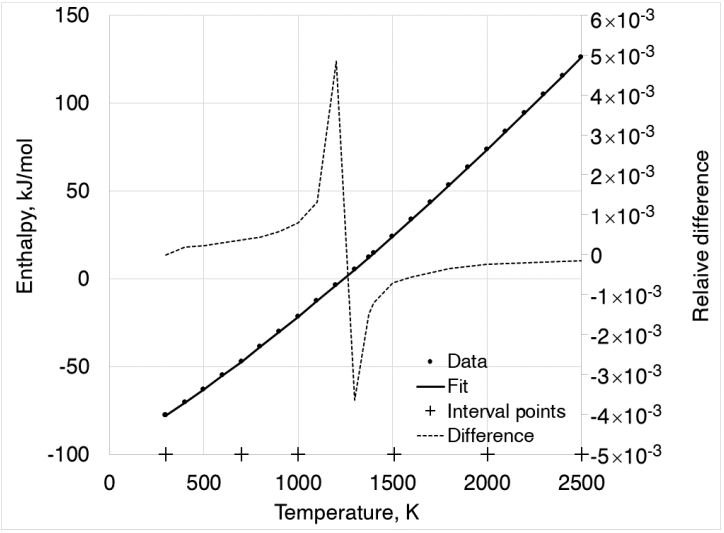


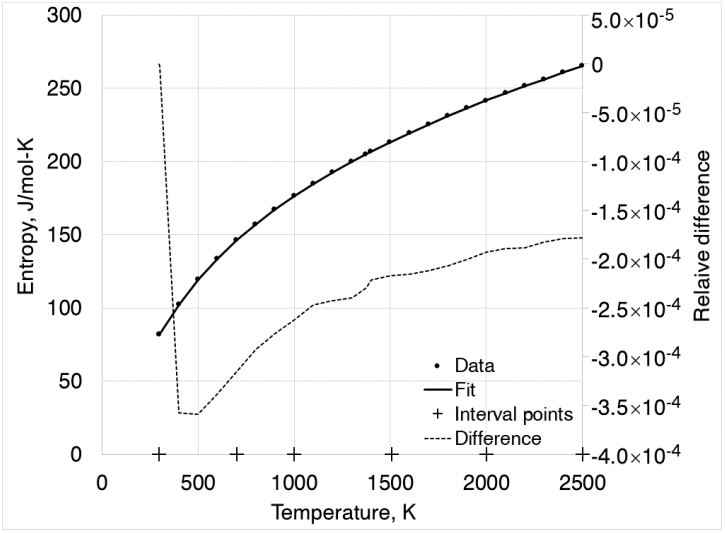


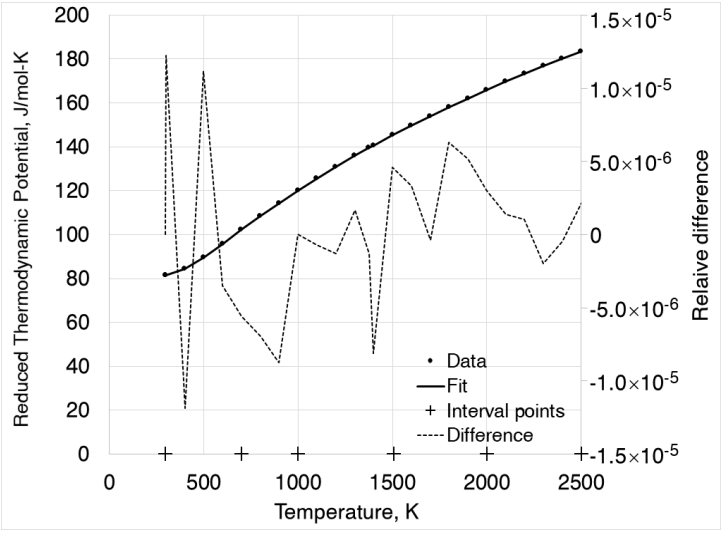
Magnesium Silicide (Mg_2Si) – Mg2Si1(cr)

MW = 76.696 g/mol – $\Delta_f H^o$(298.15 K) = -77.822 kJ/mol – S^o(298.15 K) = 81.588 J/mol-K – p^o = 0.1 MPa



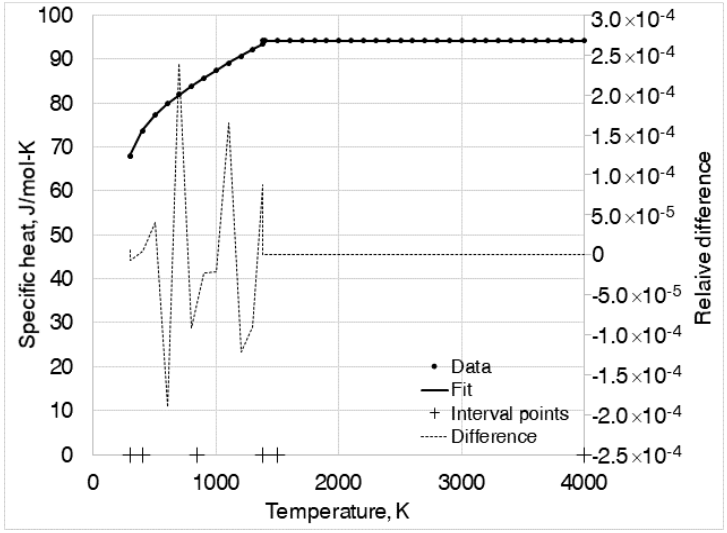


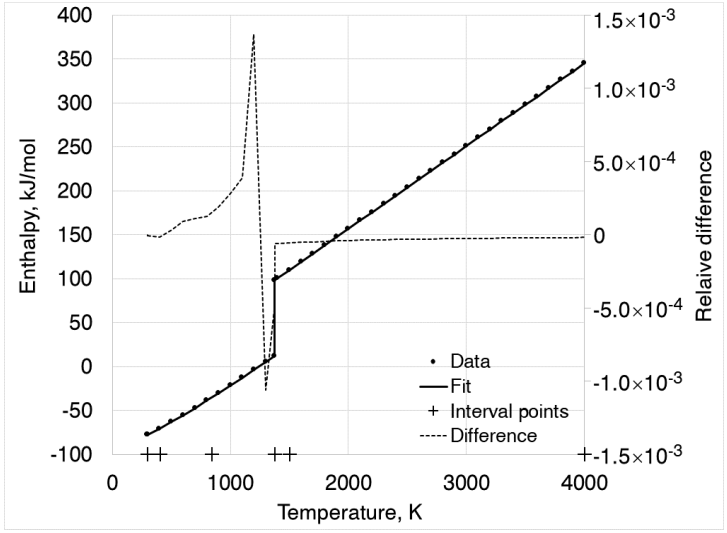


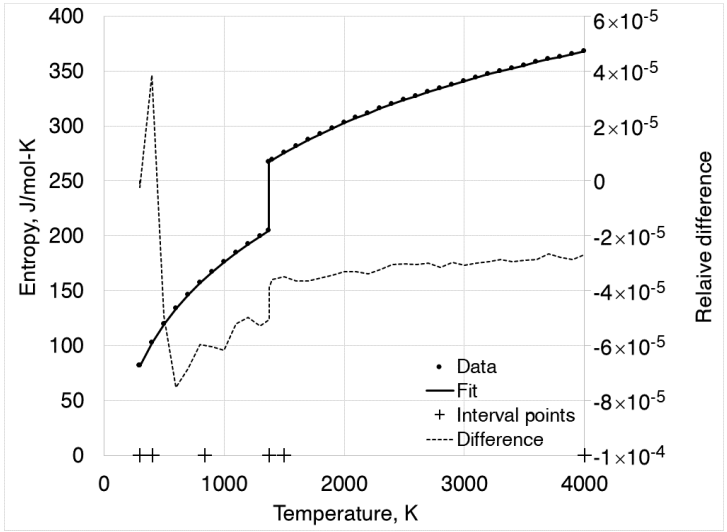


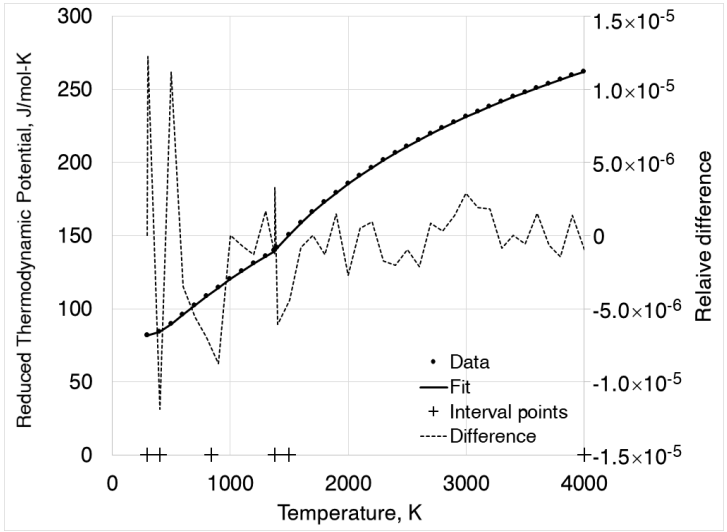
Magnesium Silicide (Mg_2Si) – Mg2Si1(cr,l)

MW = 76.696 g/mol – $\Delta_f H^o$(298.15 K) = -77.822 kJ/mol – S^o(298.15 K) = 81.588 J/mol-K – p^o = 0.1 MPa



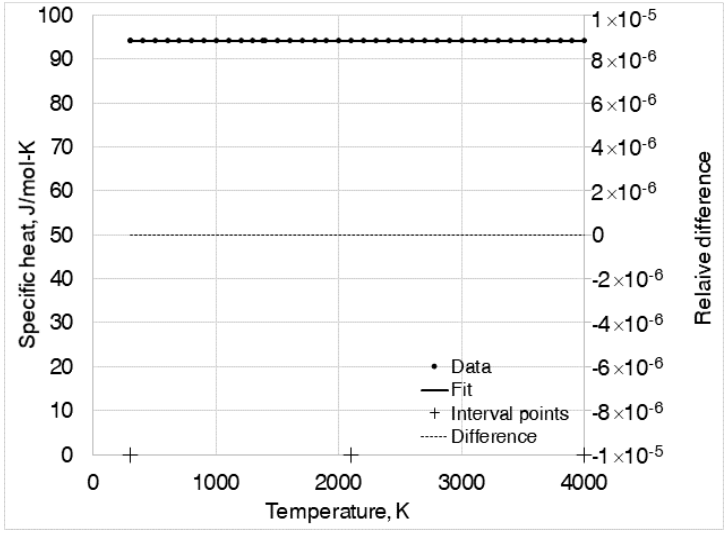


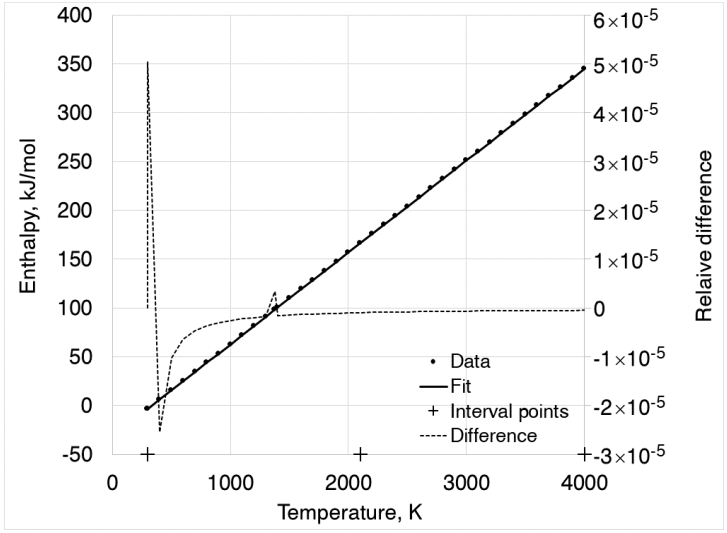


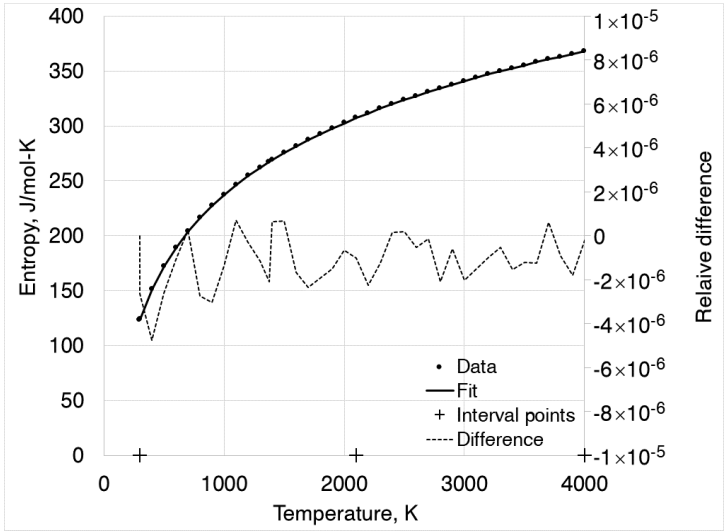


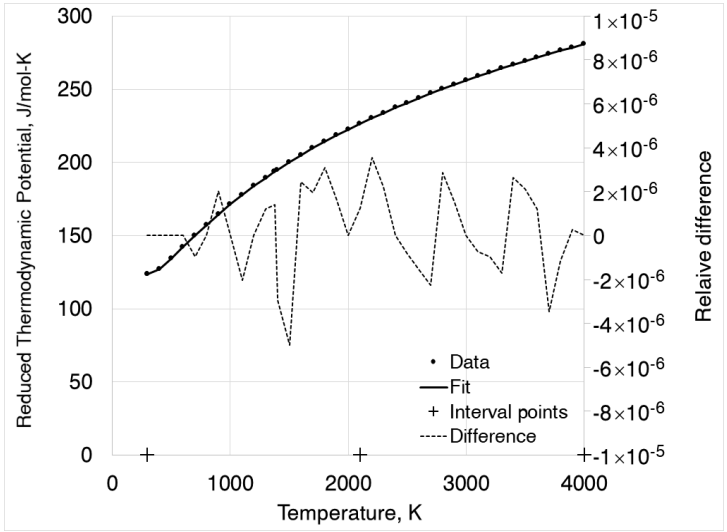
Magnesium Silicide (Mg_2Si) – Mg2Si1(l)

MW = 76.696 g/mol – $\Delta_f H^o$(298.15 K) = -3.327 kJ/mol – S^o(298.15 K) = 123.472 J/mol-K – p^o = 0.1 MPa



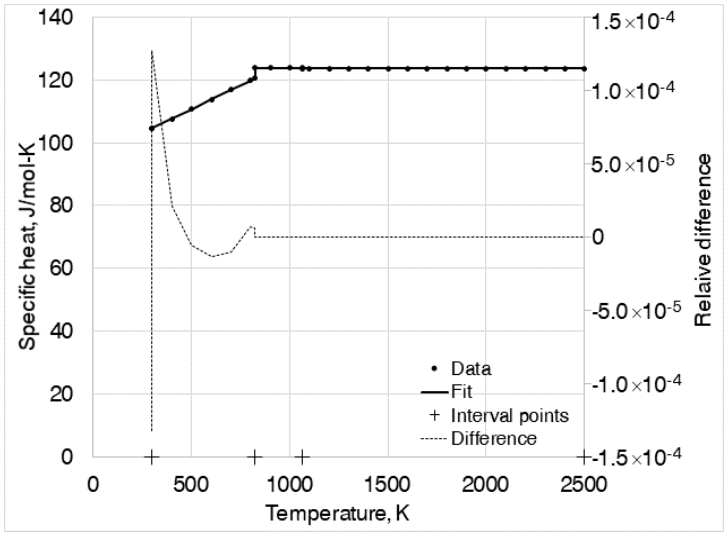


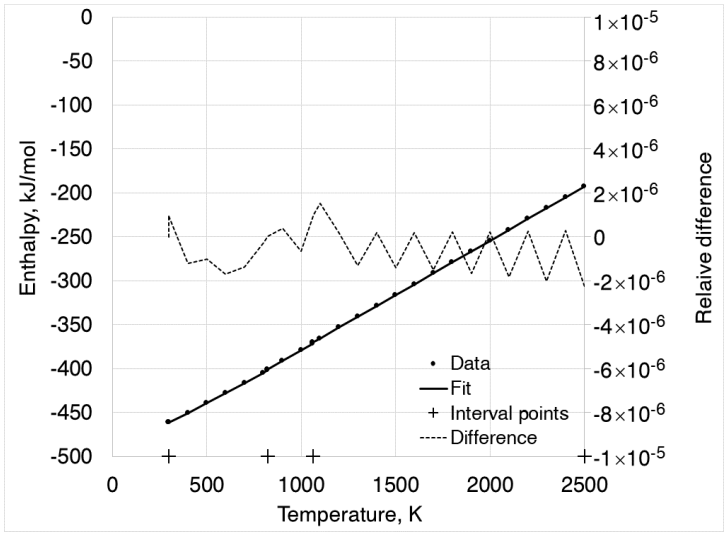


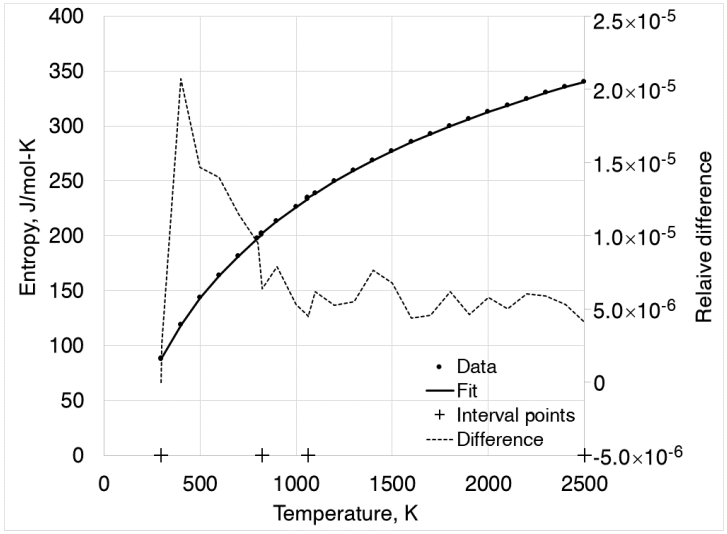


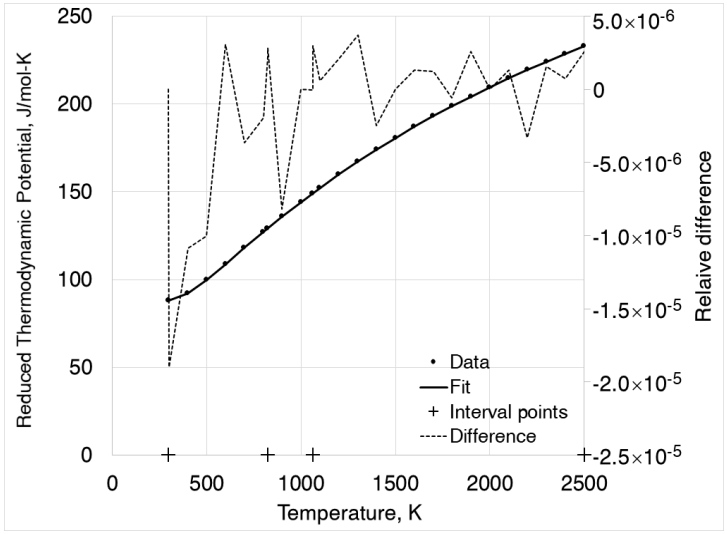
Magnesium Nitride (Mg_3N_2) – Mg3N2(cr)

MW = 100.928 g/mol – $\Delta_f H^o$(298.15 K) = -461.077 kJ/mol – S^o(298.15 K) = 87.864 J/mol-K – p^o = 0.1 MPa



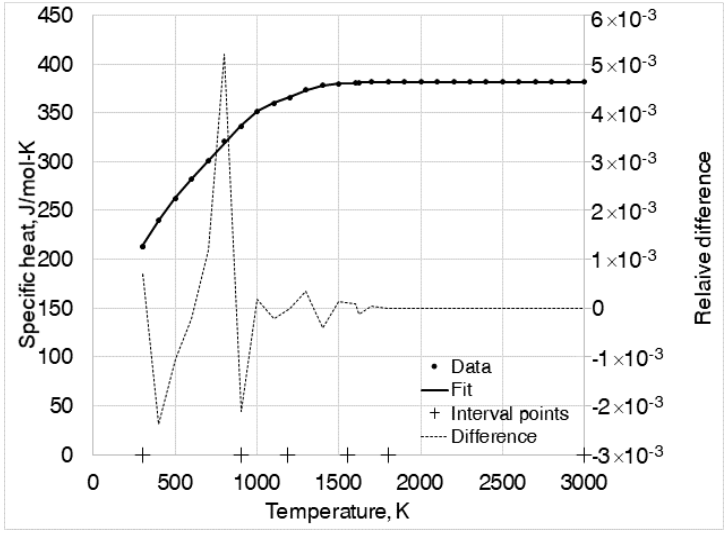


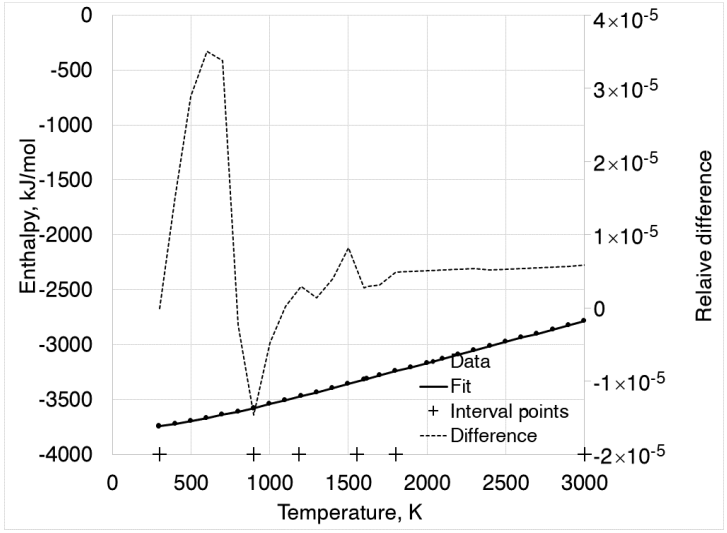


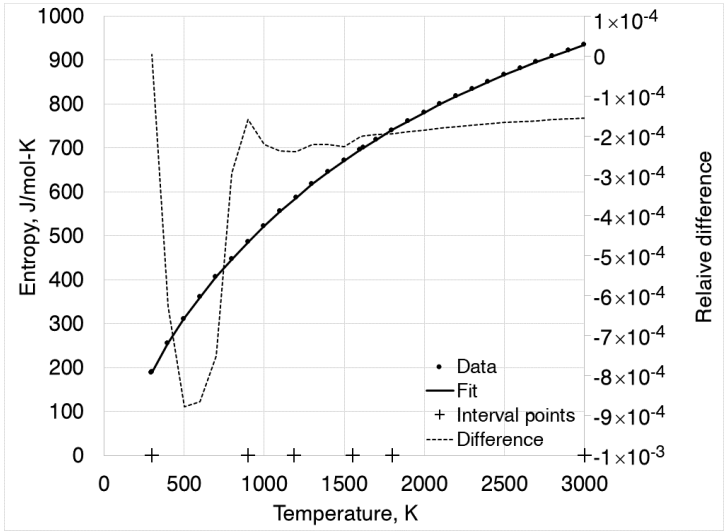


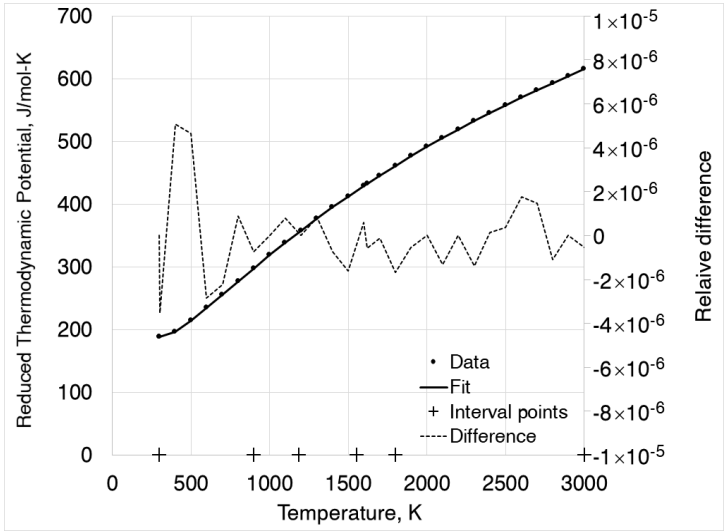
Magnesium Phosphate ($Mg_3P_2O_8$) – Mg3O8P2(cr)

MW = 262.858 g/mol – $\Delta_f H^o$(298.15 K) = -3745.098 kJ/mol – S^o(298.15 K) = 188.28 J/mol-K – p^o = 0.1 MPa



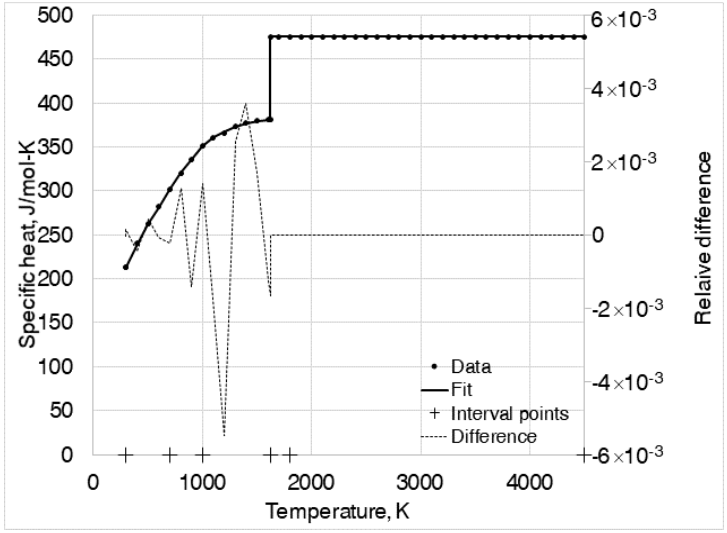


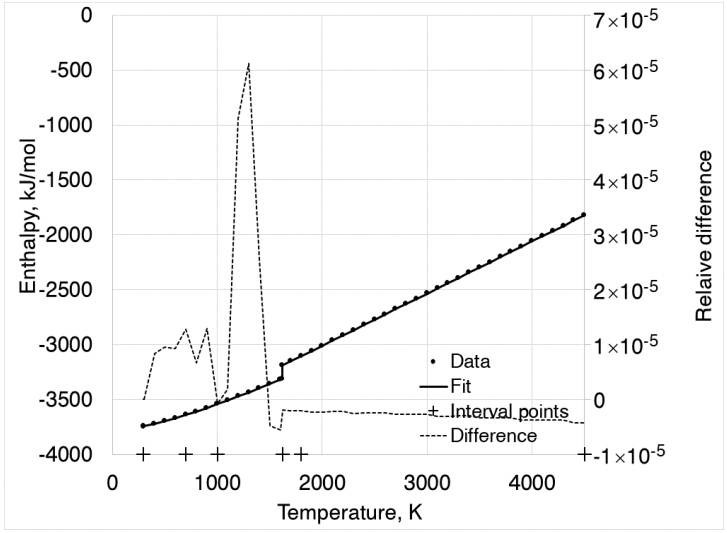


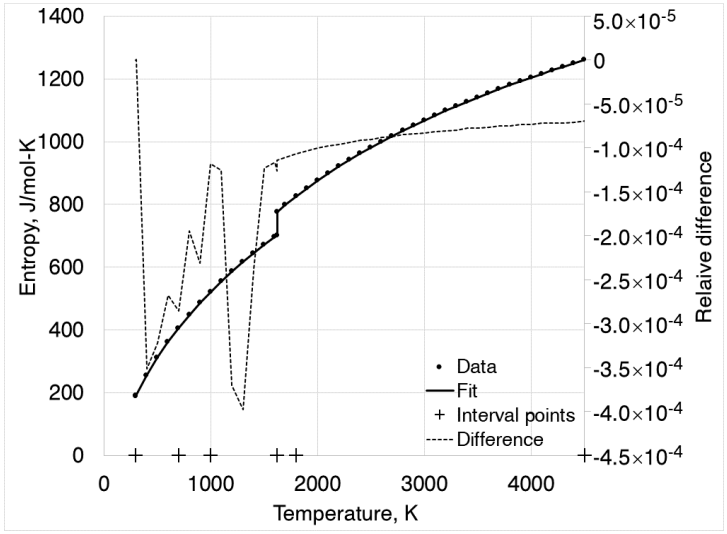


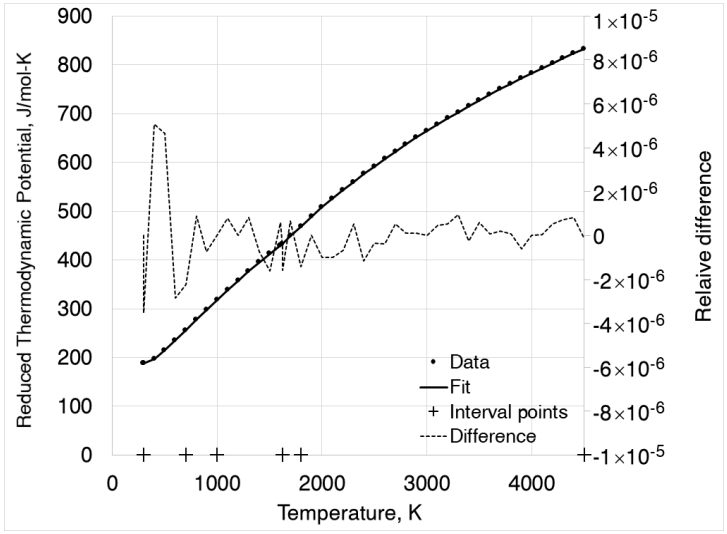
Magnesium Phosphate ($Mg_3P_2O_8$) – Mg3O8P2(cr,l)

MW = 262.858 g/mol – $\Delta_f H^o$(298.15 K) = -3745.098 kJ/mol – S^o(298.15 K) = 188.28 J/mol-K – p^o = 0.1 MPa



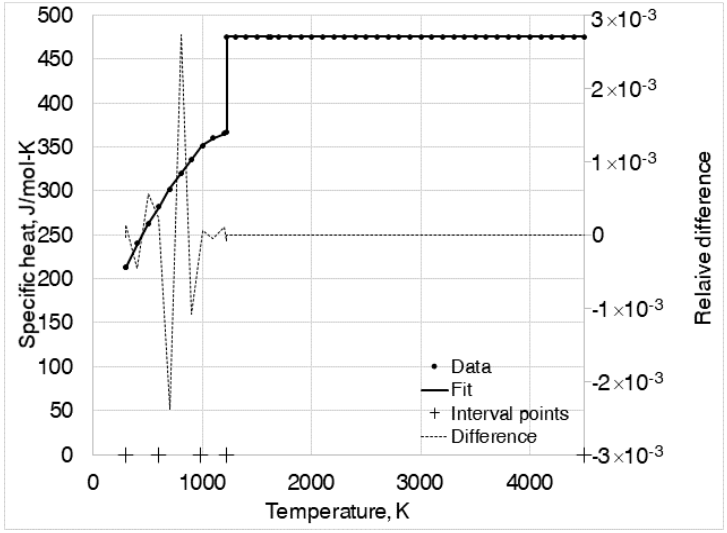


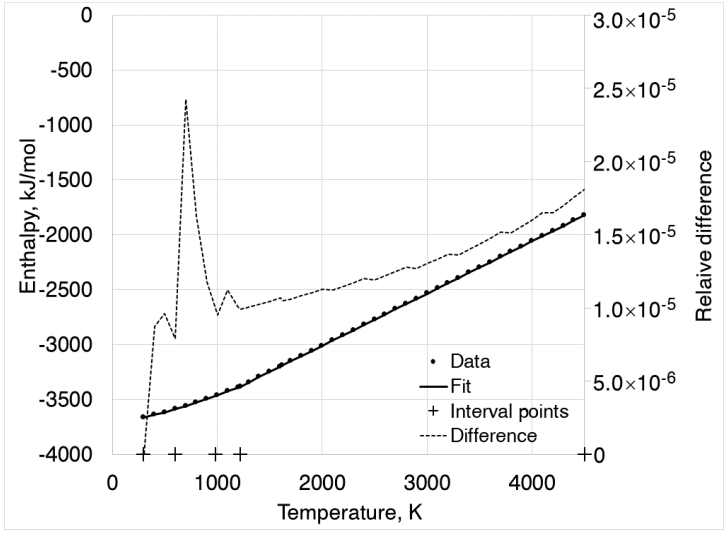


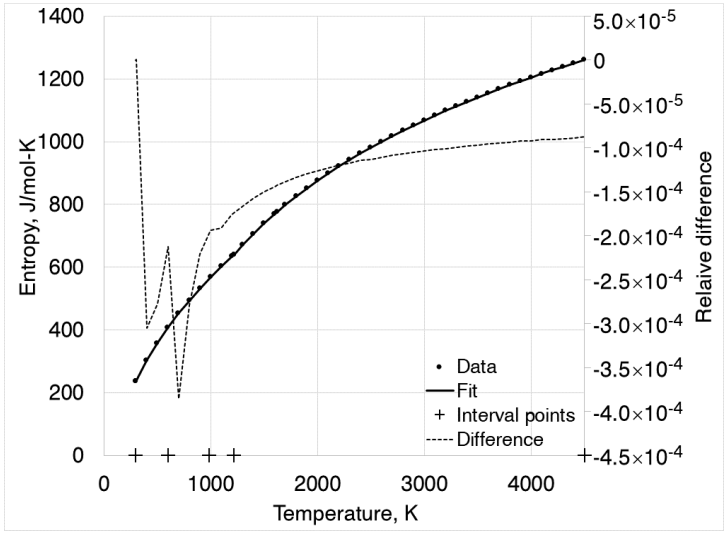


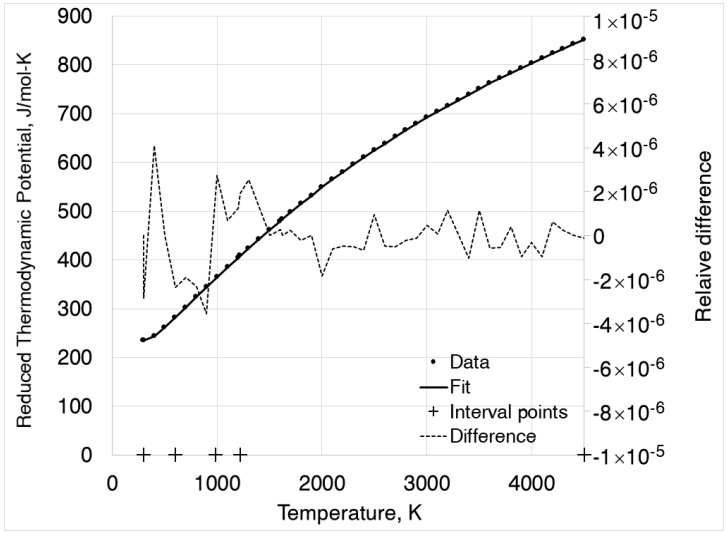
Magnesium Phosphate ($Mg_3P_2O_8$) – Mg3O8P2(l)

MW = 262.858 g/mol – $\Delta_f H^o$(298.15 K) = -3663.19 kJ/mol – S^o(298.15 K) = 235.102 J/mol-K – p^o = 0.1 MPa



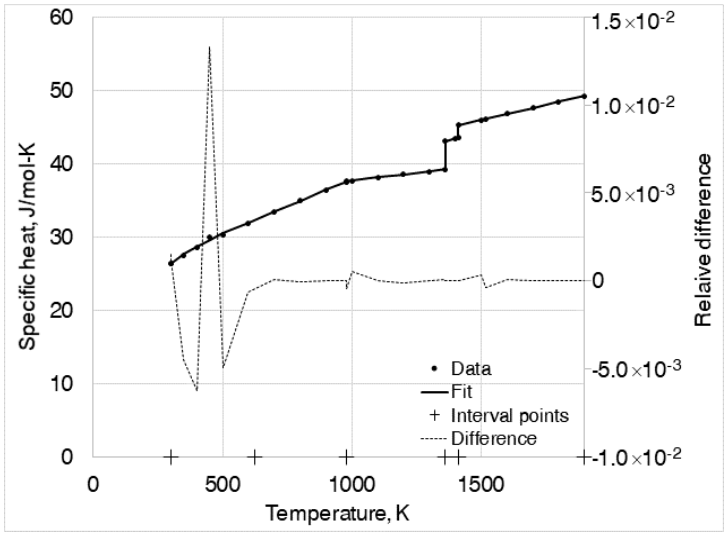


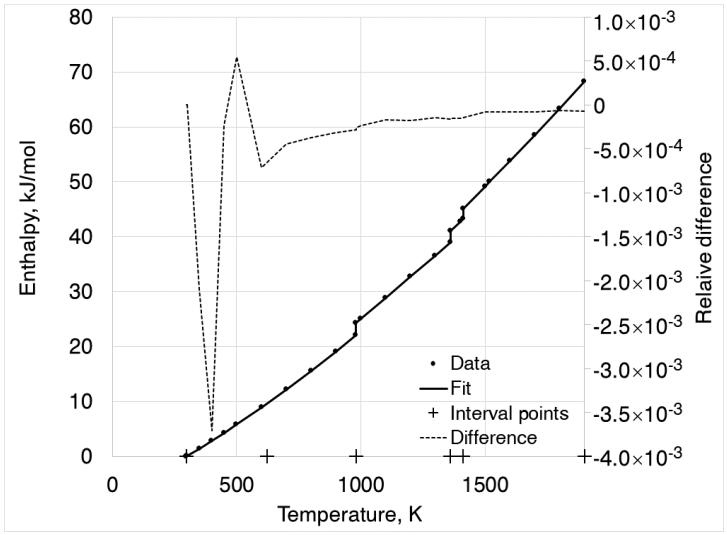


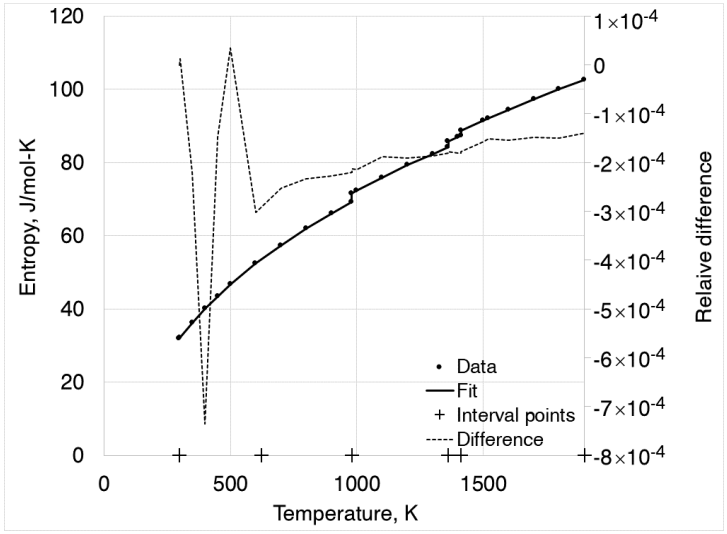


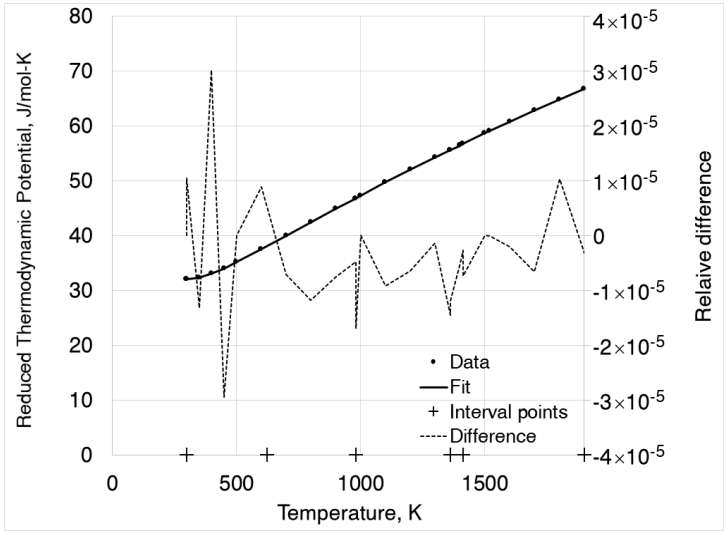
Manganese (Mn) – Mn1(cr)

MW = 54.938 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 32.01 J/mol-K – p^o = 0.1 MPa



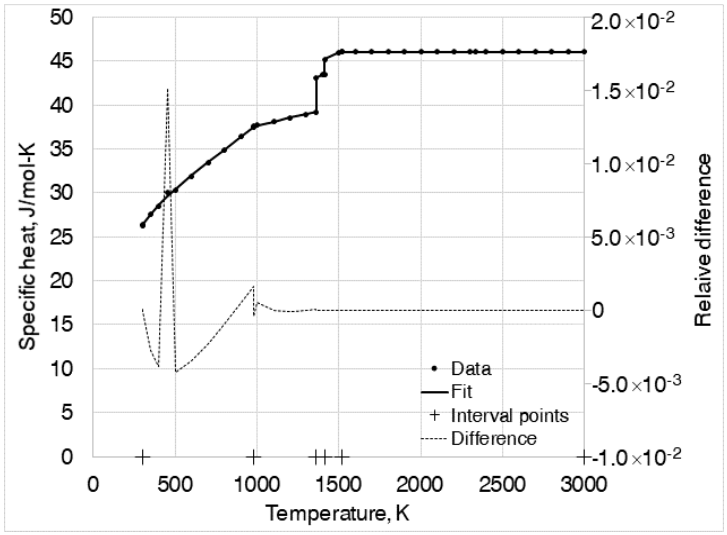


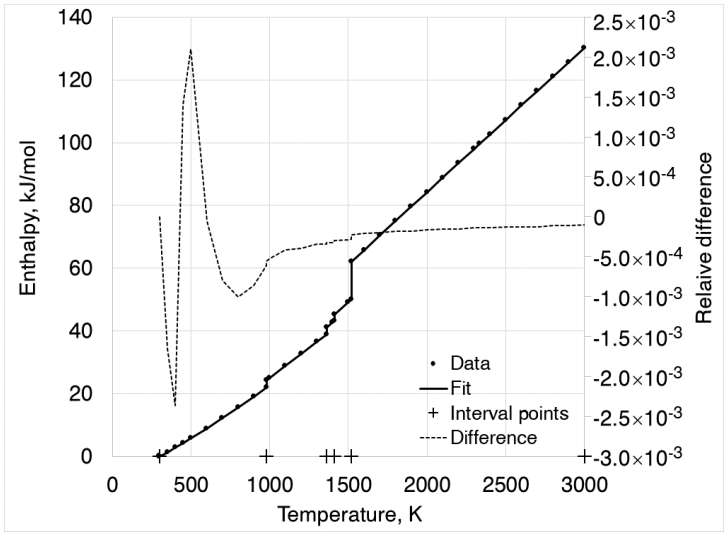


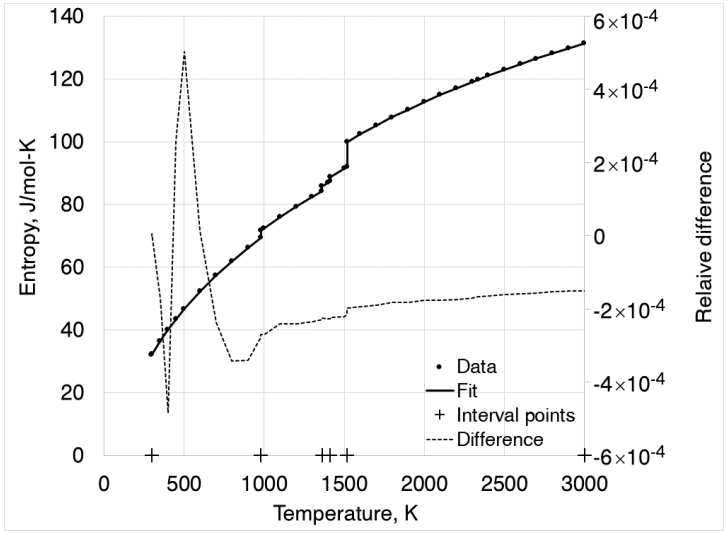


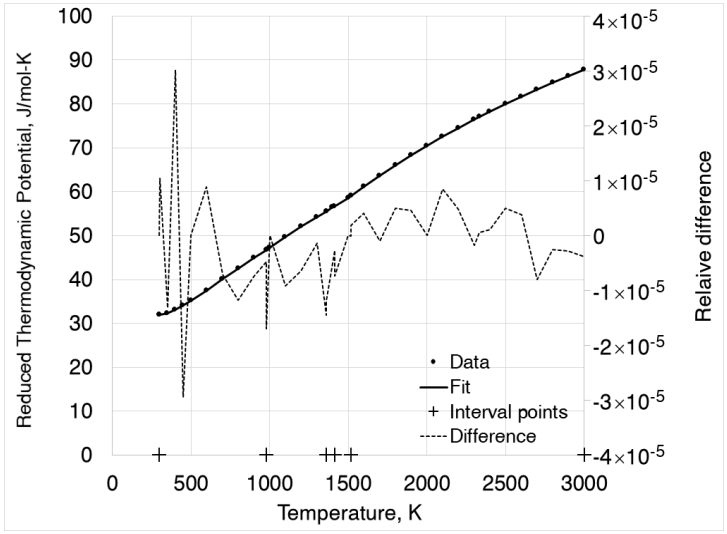
Manganese (Mn) – Mn1(cr,l)

MW = 54.938 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 32.01 J/mol-K – p^o = 0.1 MPa



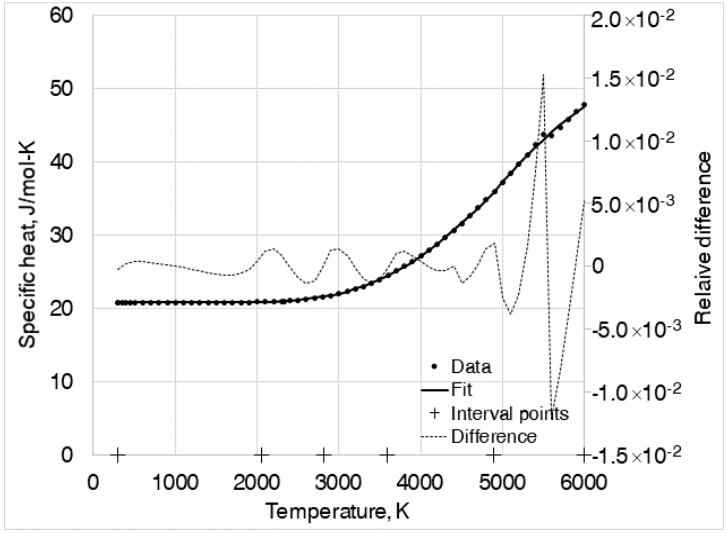


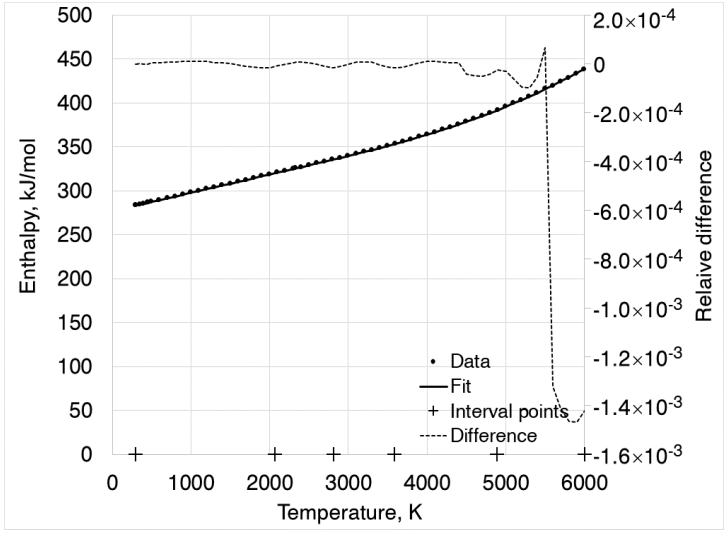


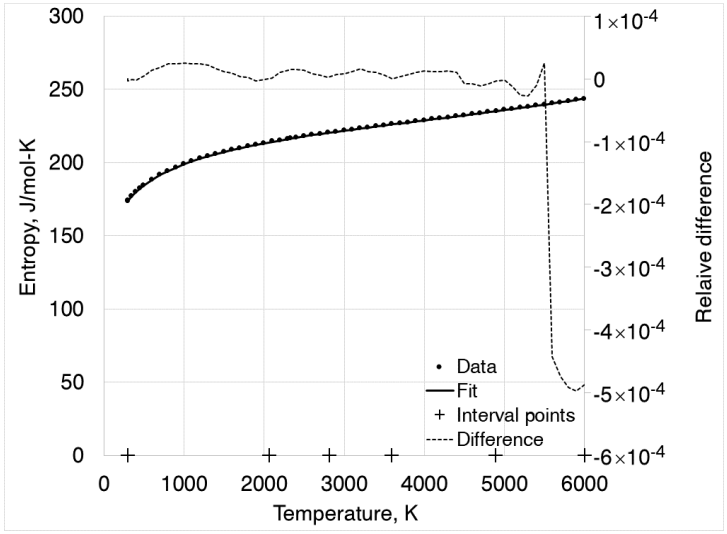


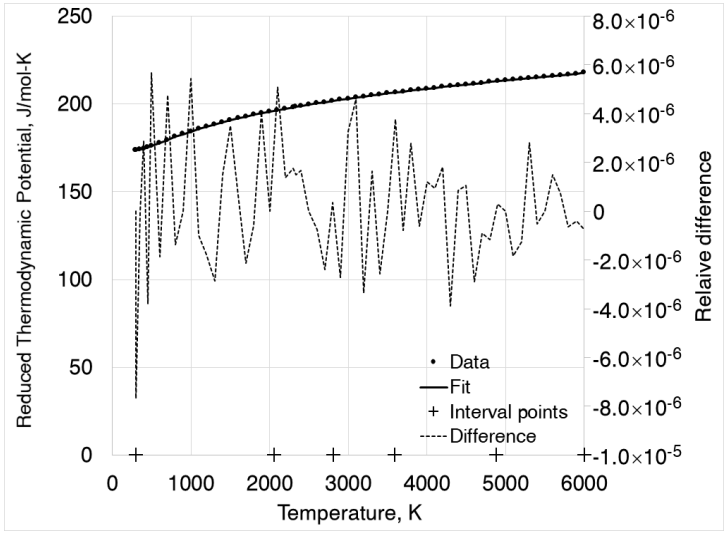
Manganese (Mn) – Mn1(g)

MW = 54.938 g/mol – $\Delta_f H^o$(298.15 K) = 283.257 kJ/mol – S^o(298.15 K) = 173.716 J/mol-K – p^o = 0.1 MPa



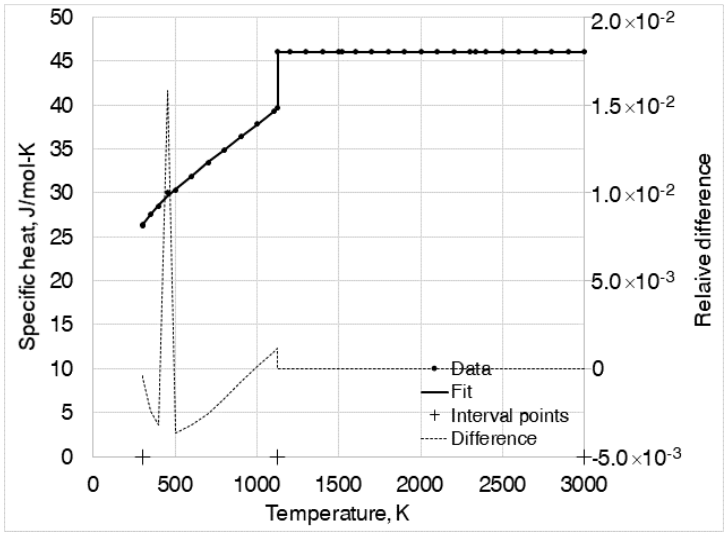


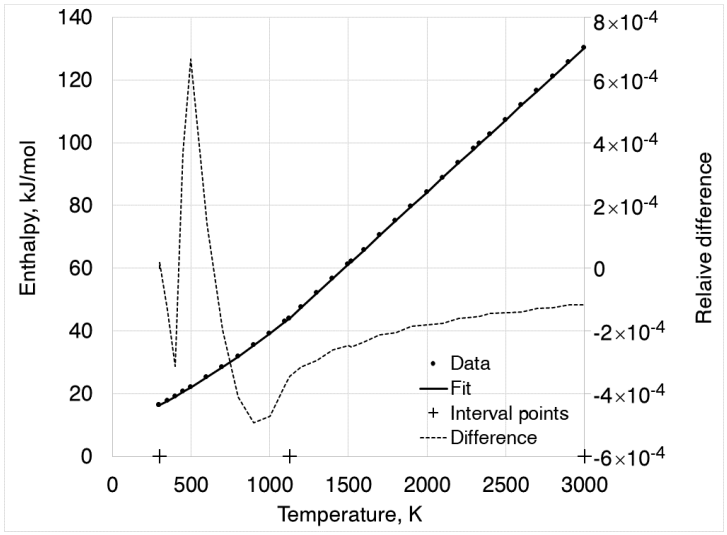


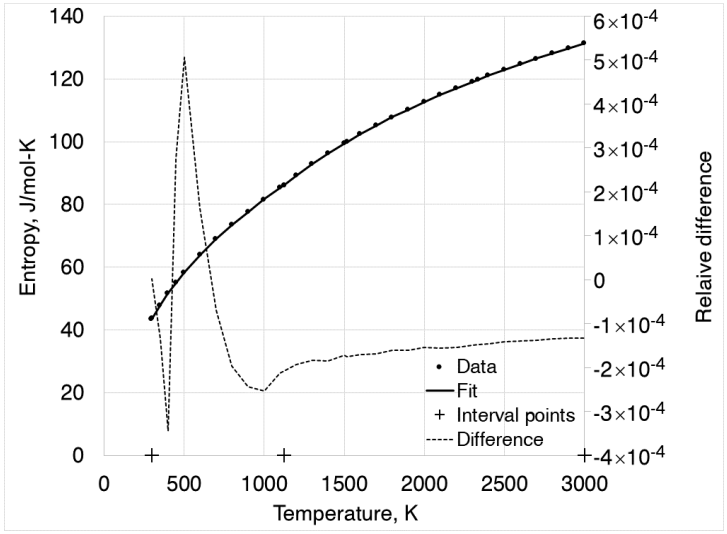


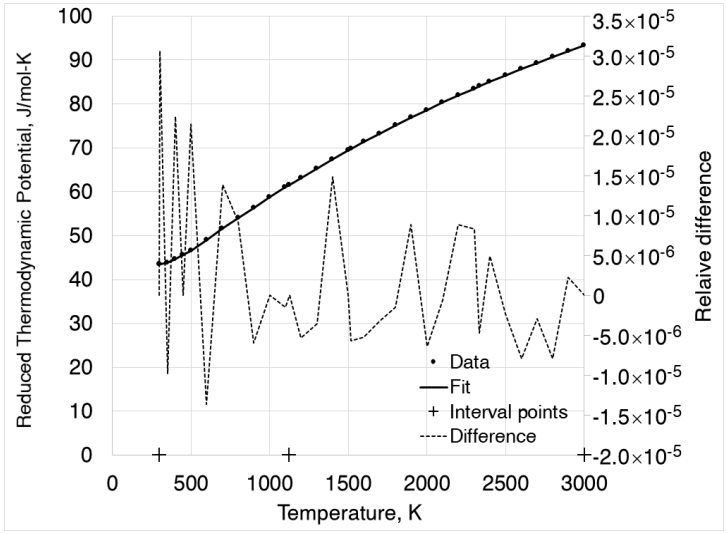
Manganese (Mn) – Mn1(l)

MW = 54.938 g/mol – $\Delta_f H^o$(298.15 K) = 16.289 kJ/mol – S^o(298.15 K) = 43.499 J/mol-K – p^o = 0.1 MPa



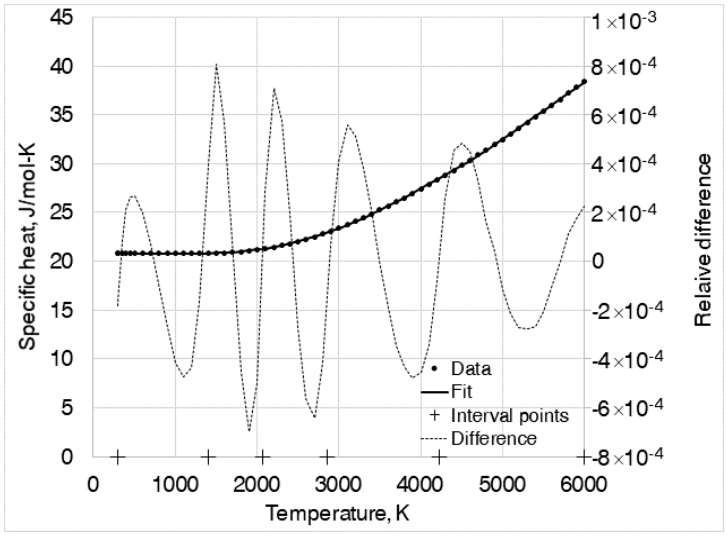


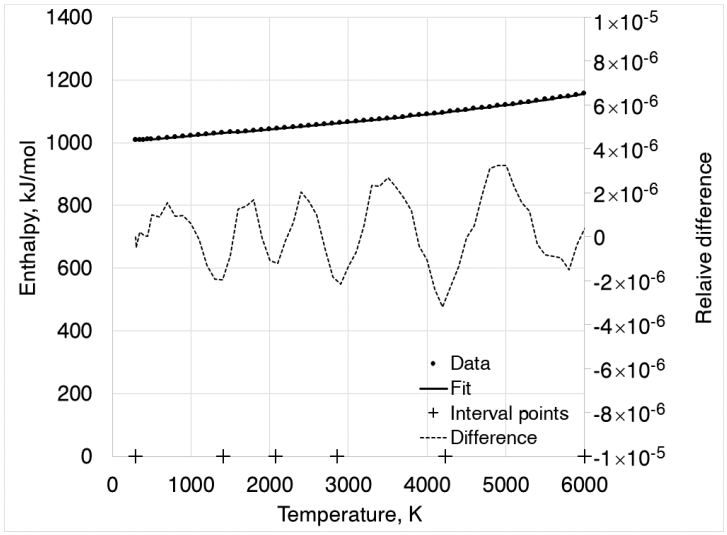


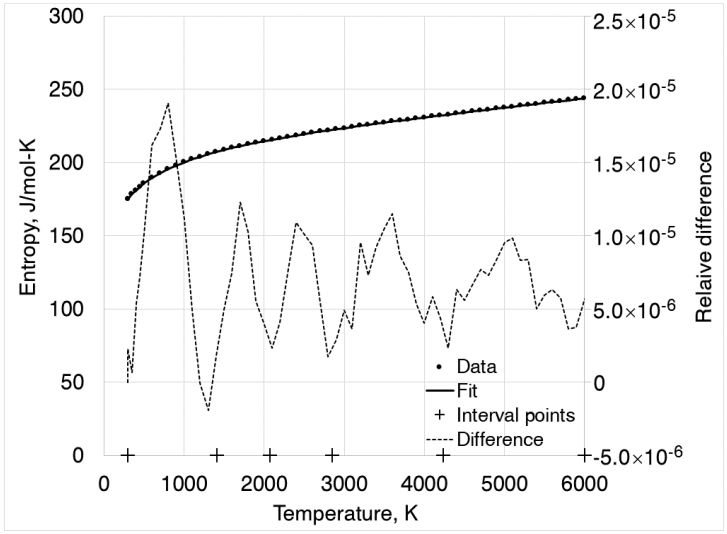


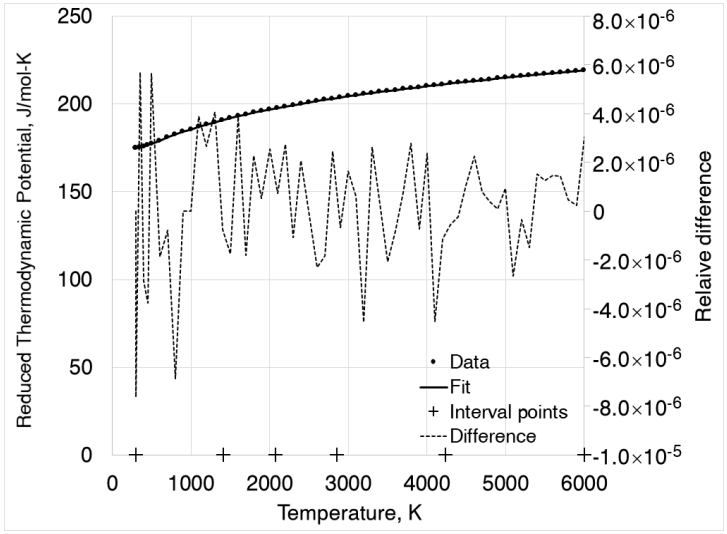
Manganese, Ion (Mn^+) – Mn1+(g)

MW = 54.9375 g/mol – $\Delta_f H^o$(298.15 K) = 1006.986 kJ/mol – S^o(298.15 K) = 174.997 J/mol-K – p^o = 0.1 MPa



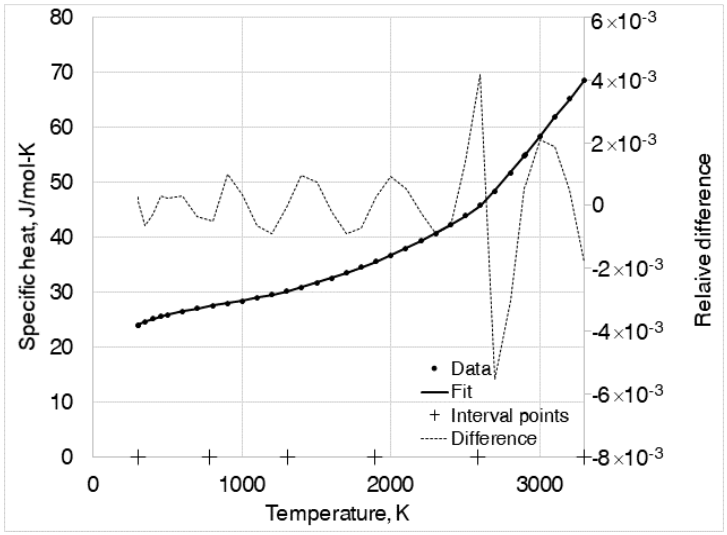


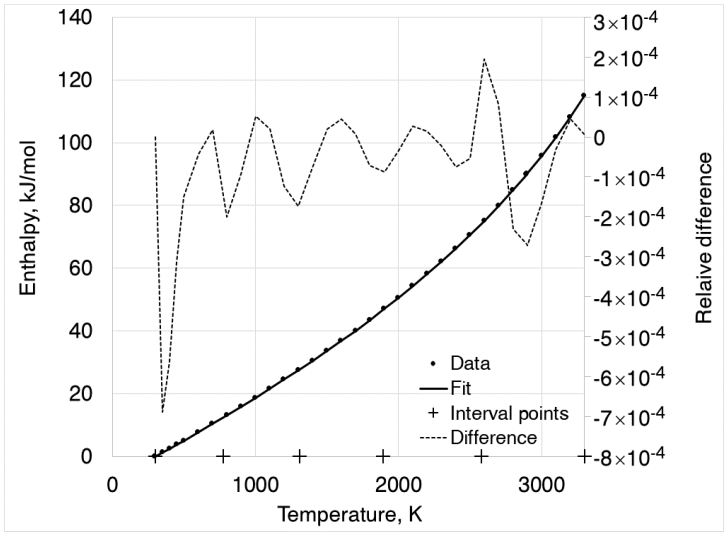


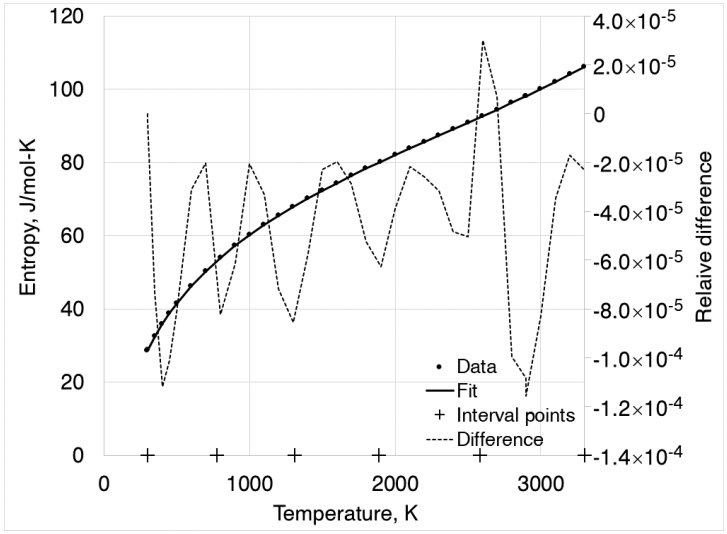


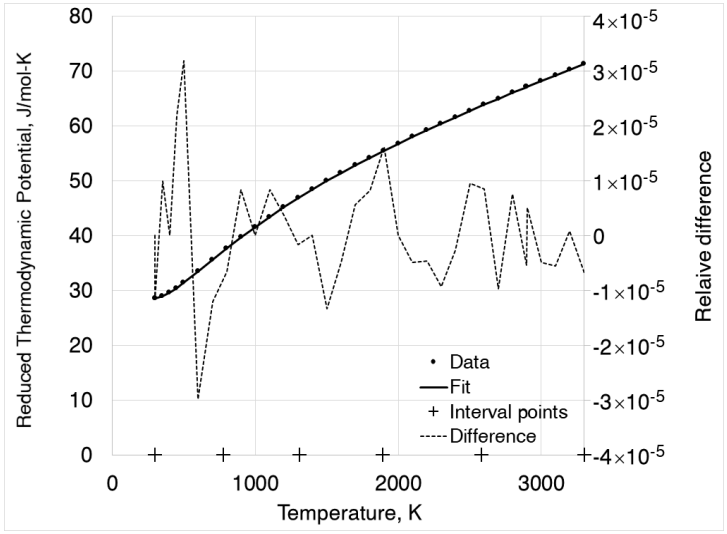
Molybdenum (Mo) – Mo1(cr)

MW = 95.94 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 28.605 J/mol-K – p^o = 0.1 MPa



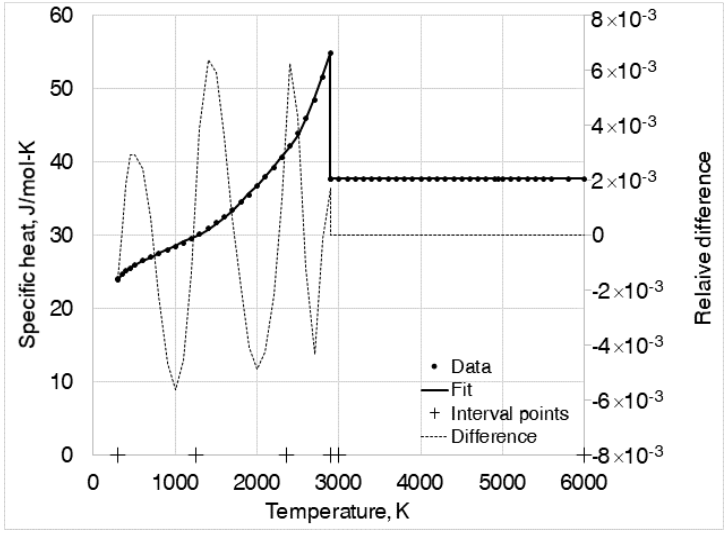


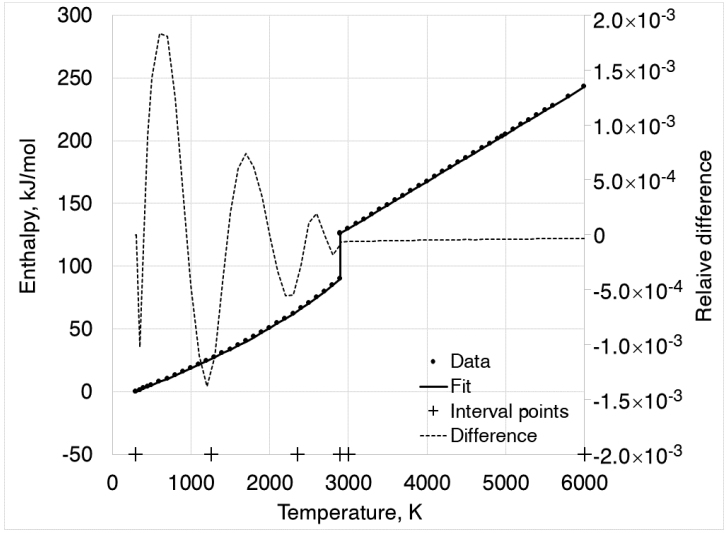


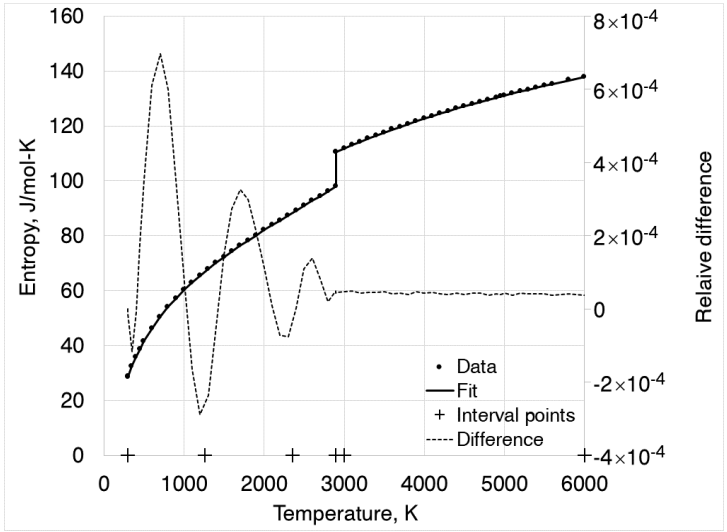


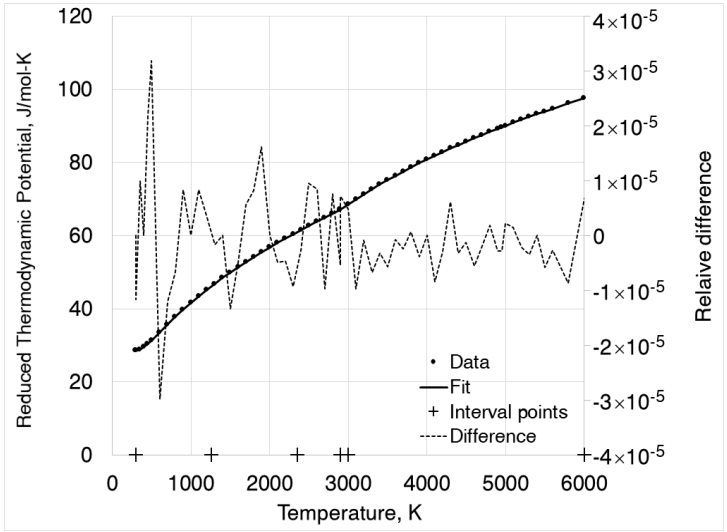
Molybdenum (Mo) – Mo1(cr,l)

MW = 95.94 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 28.605 J/mol-K – p^o = 0.1 MPa



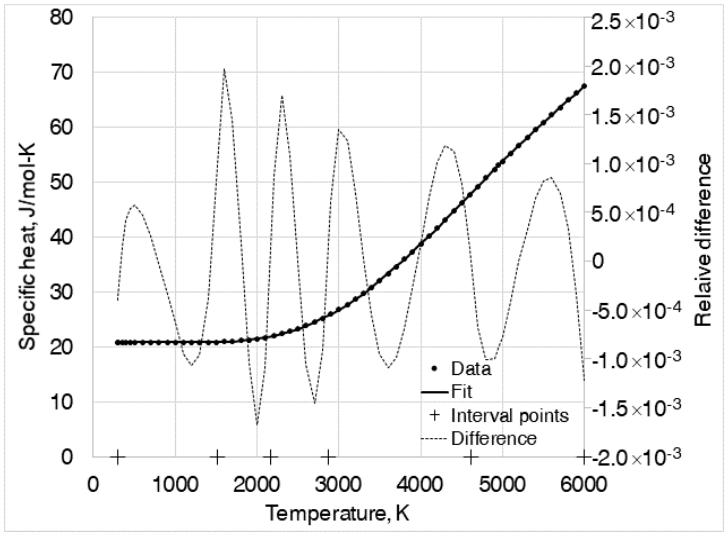


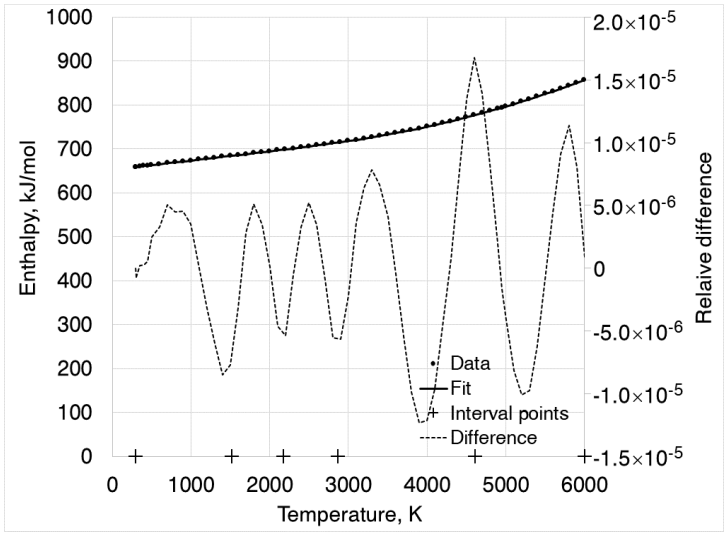


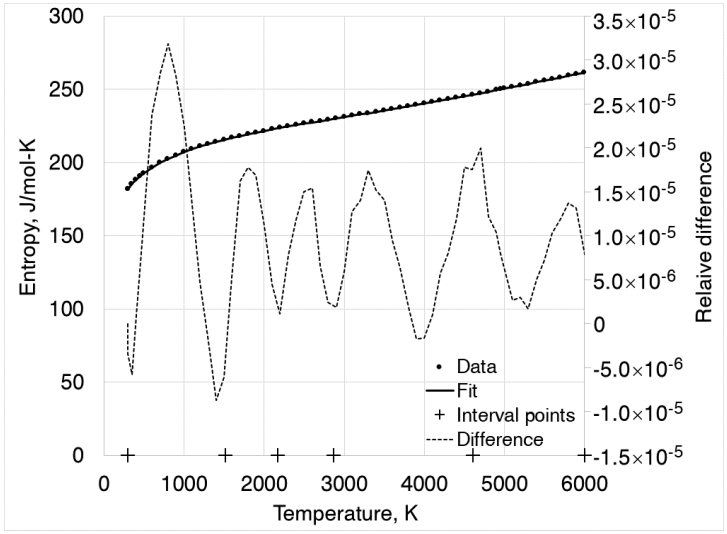


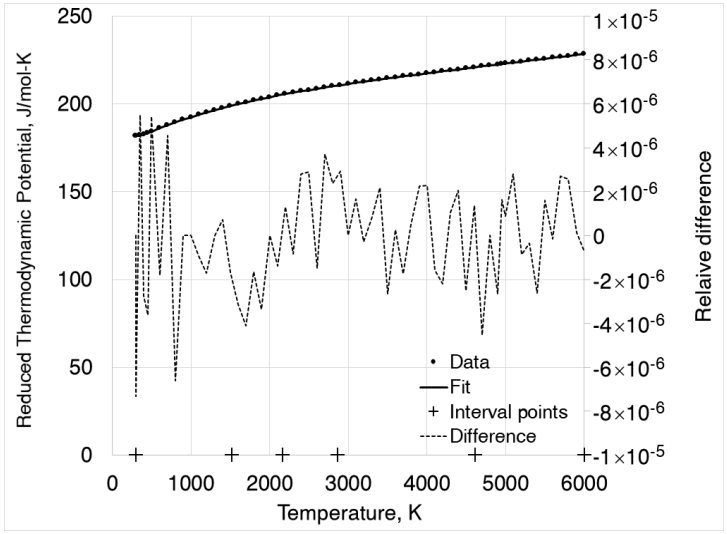
Molybdenum (Mo) – Mo1(g)

MW = 95.94 g/mol – $\Delta_f H^o$(298.15 K) = 658.98 kJ/mol – S^o(298.15 K) = 181.951 J/mol-K – p^o = 0.1 MPa



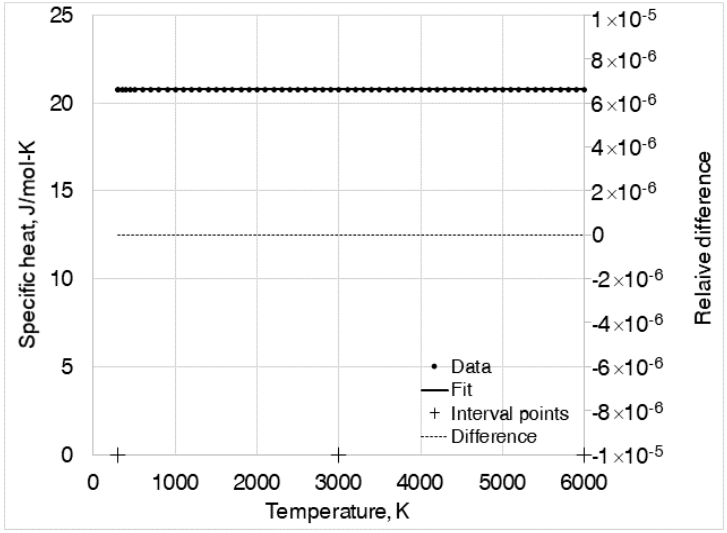


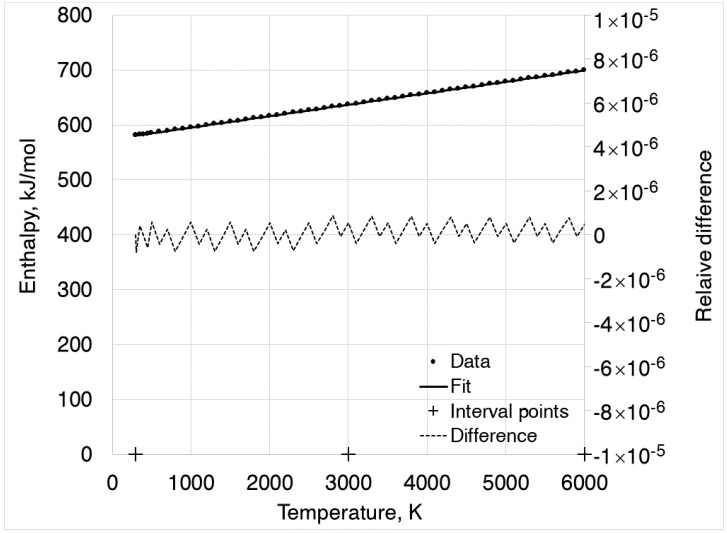


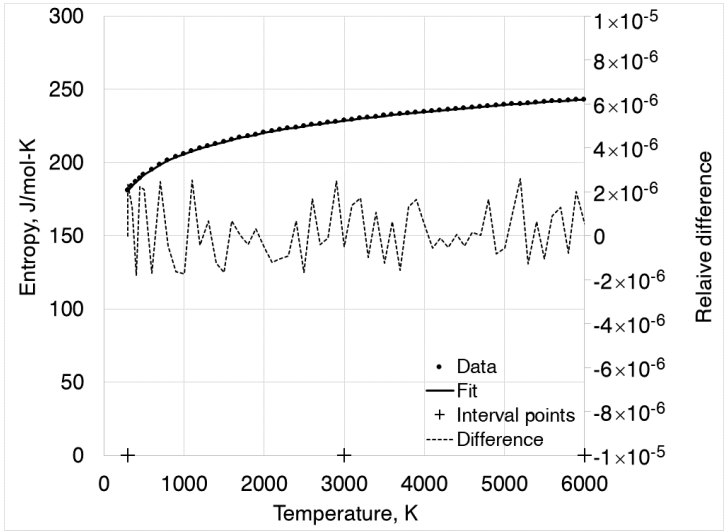


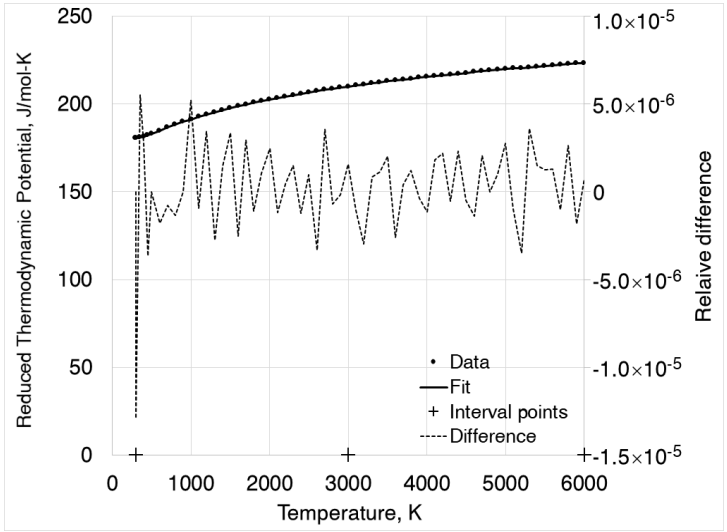
Molybdenum, Ion (Mo^-) – Mo1-(g)

MW = 95.94 g/mol – $\Delta_f H^o$(298.15 K) = 580.805 kJ/mol – S^o(298.15 K) = 180.669 J/mol-K – p^o = 0.1 MPa



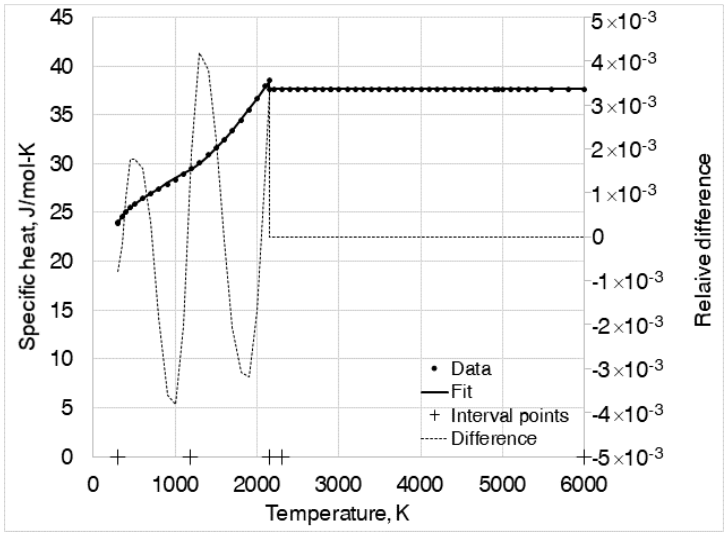


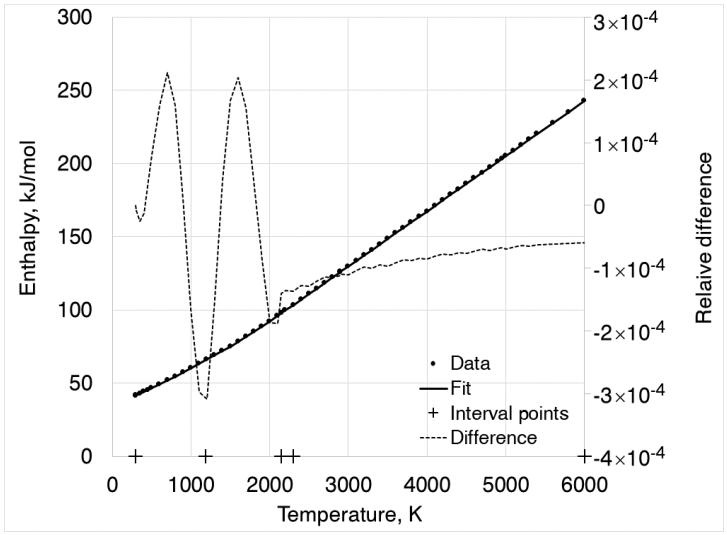


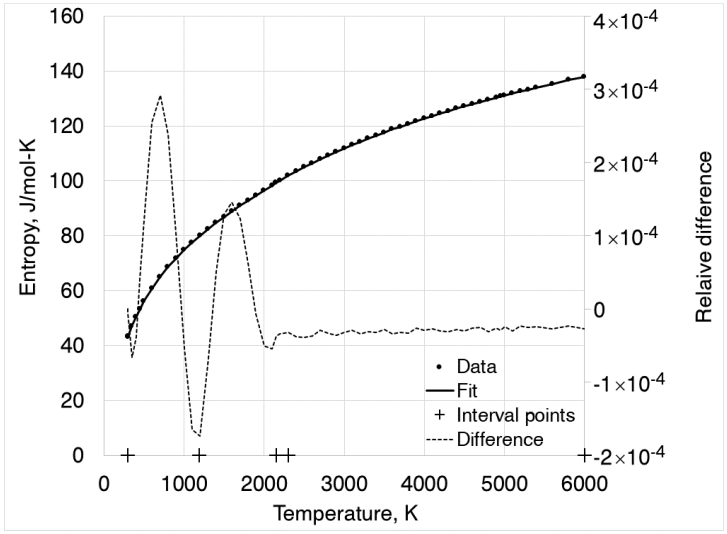


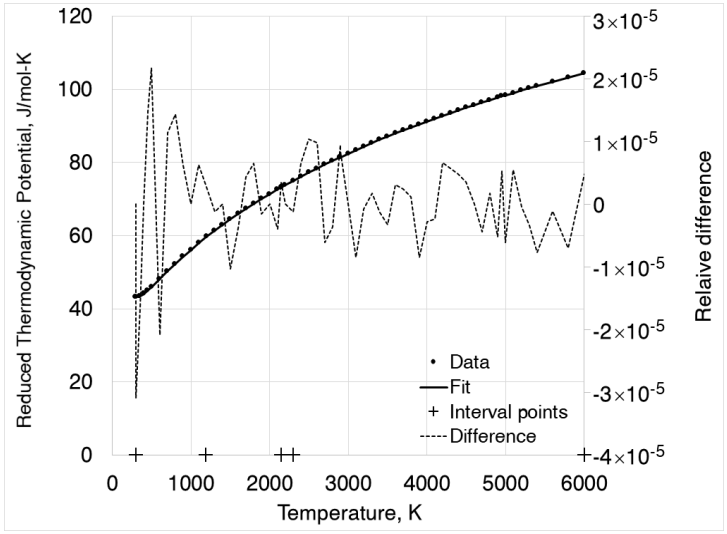
Molybdenum (Mo) – Mo1(l)

MW = 95.94 g/mol – $\Delta_f H^o$(298.15 K) = 41.568 kJ/mol – S^o(298.15 K) = 43.146 J/mol-K – p^o = 0.1 MPa



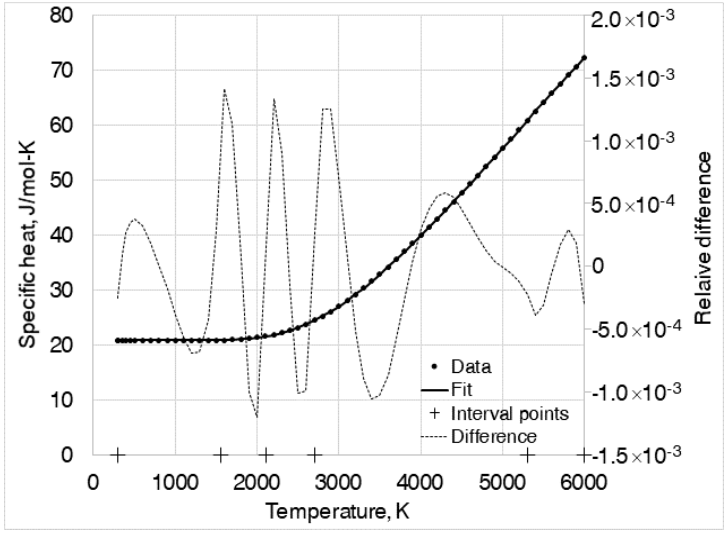


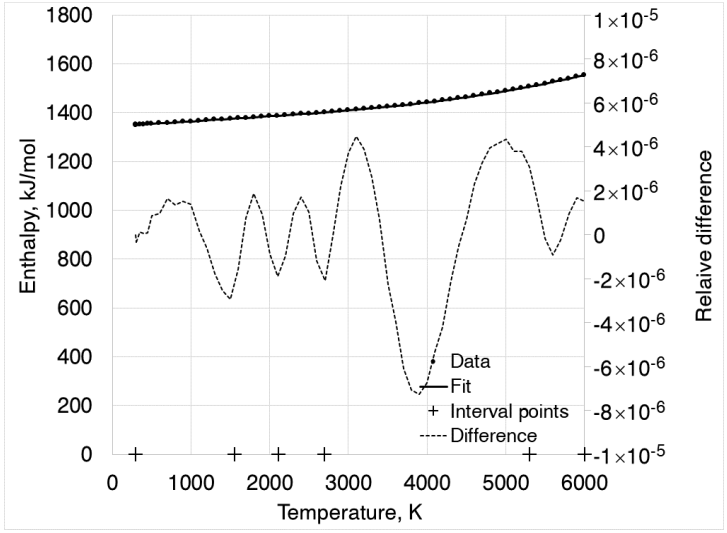


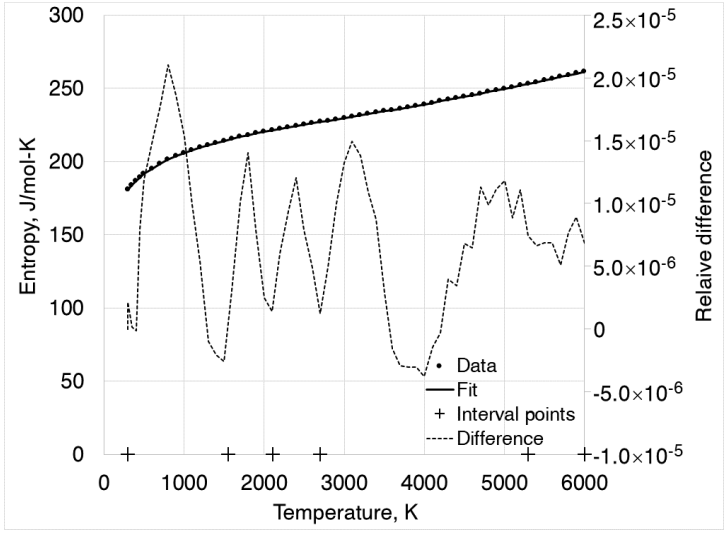


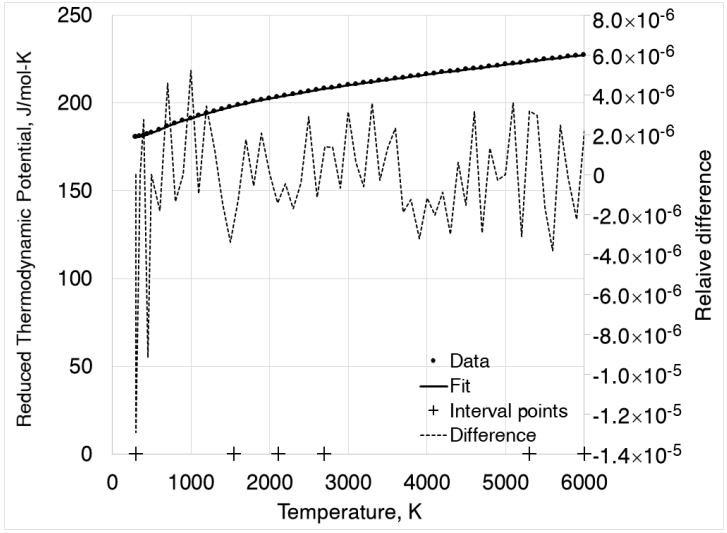
Molybdenum, Ion (Mo^+) – Mo1+(g)

MW = 95.94 g/mol – $\Delta_f H^o$(298.15 K) = 1350.159 kJ/mol – S^o(298.15 K) = 180.669 J/mol-K – p^o = 0.1 MPa



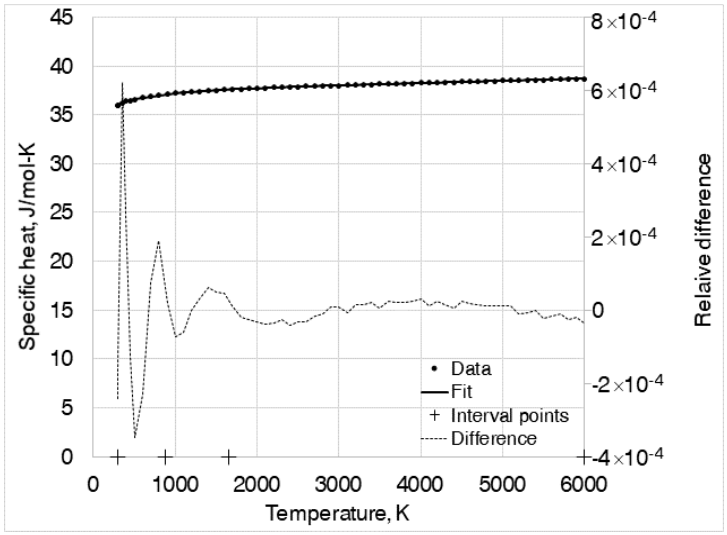


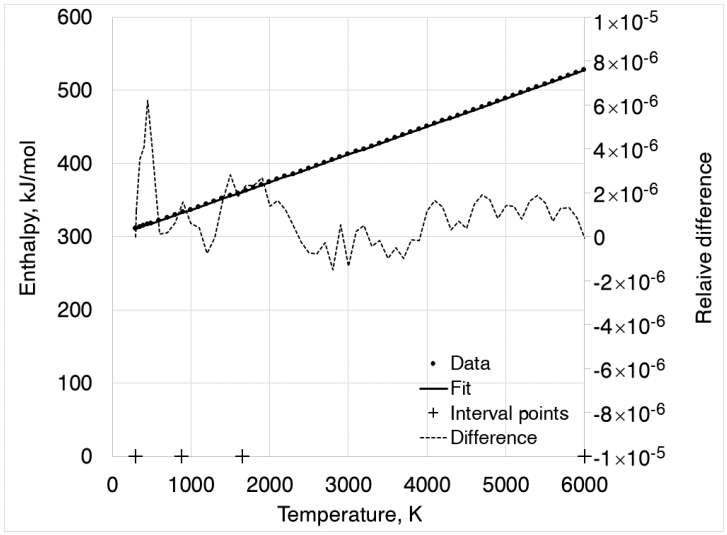


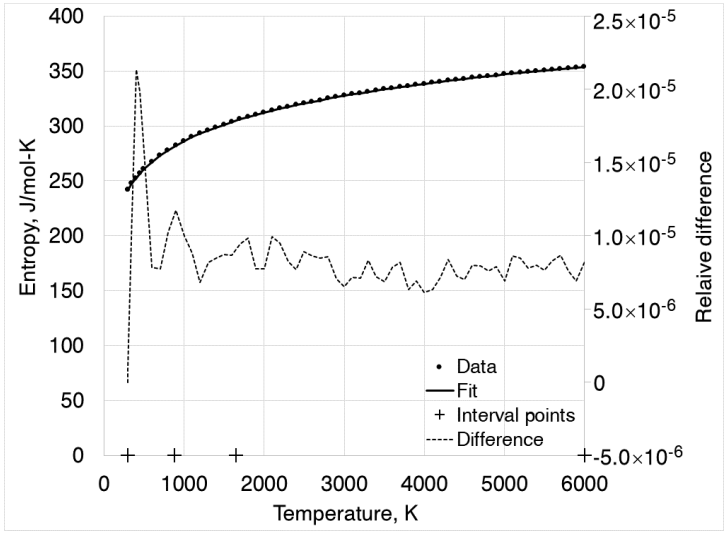


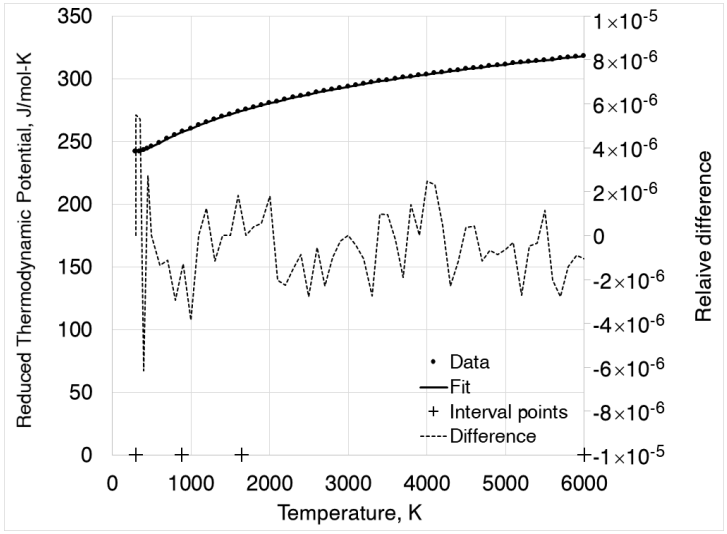
Molybdenum Oxide (MoO) – Mo1O1(g)

MW = 111.94 g/mol – $\Delta_f H^o$(298.15 K) = 310.959 kJ/mol – S^o(298.15 K) = 241.764 J/mol-K – p^o = 0.1 MPa



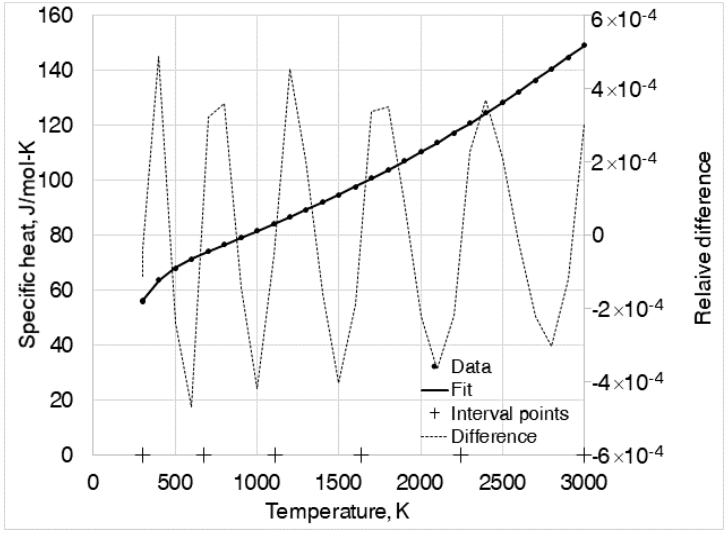


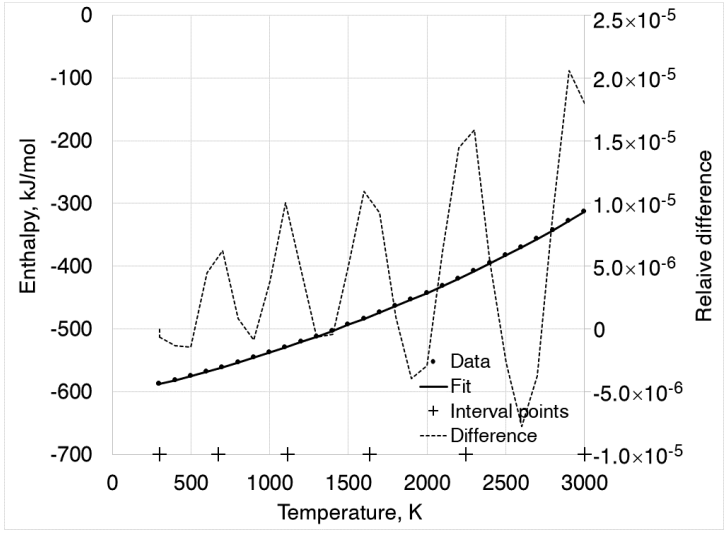


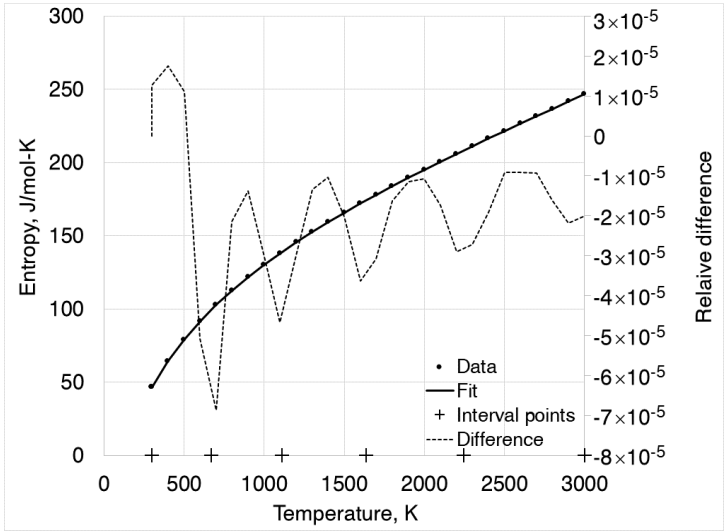


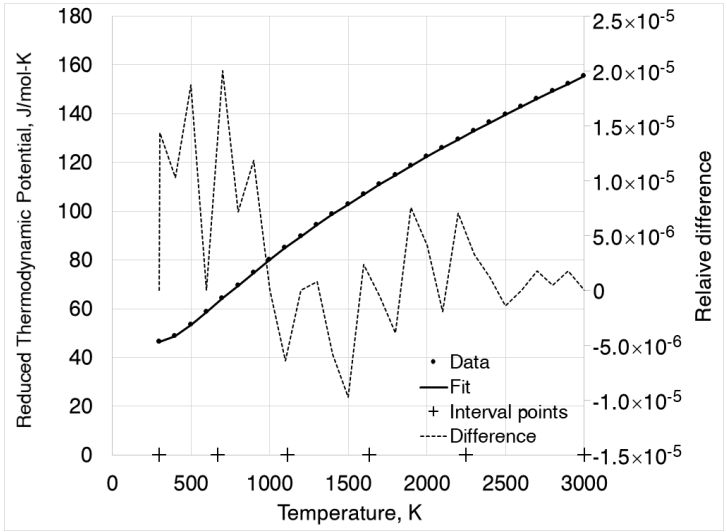
Molybdenum Oxide (MoO_2) – Mo1O2(cr)

MW = 127.94 g/mol – $\Delta_f H^o$(298.15 K) = -587.852 kJ/mol – S^o(298.15 K) = 46.459 J/mol-K – p^o = 0.1 MPa



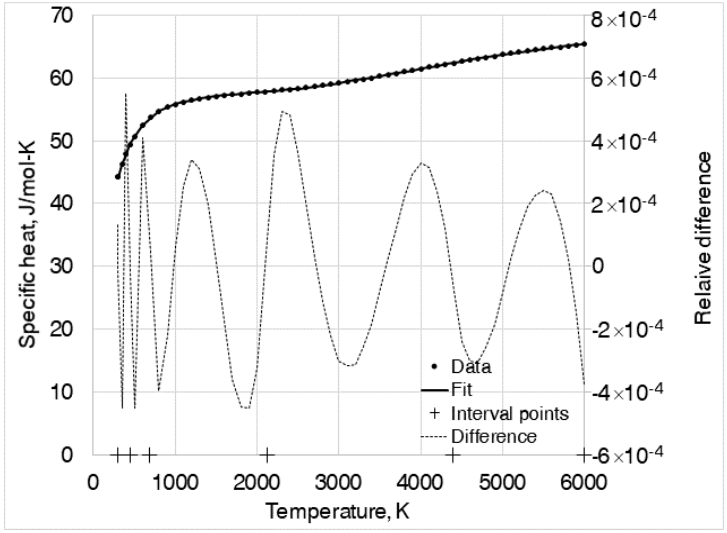


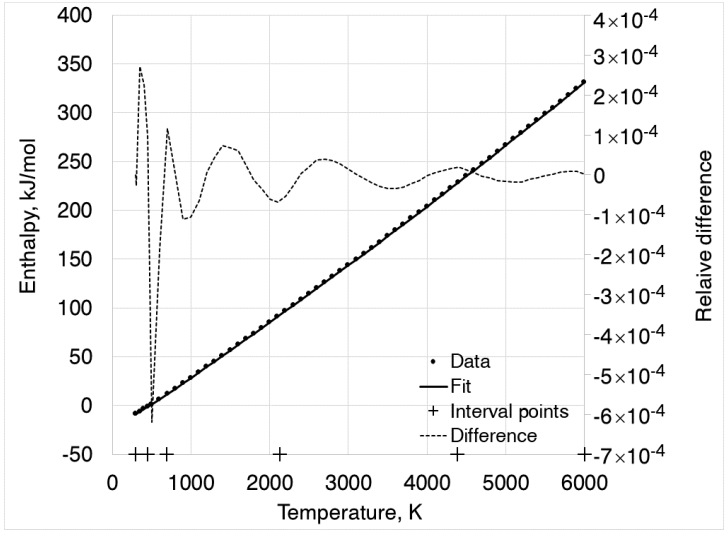


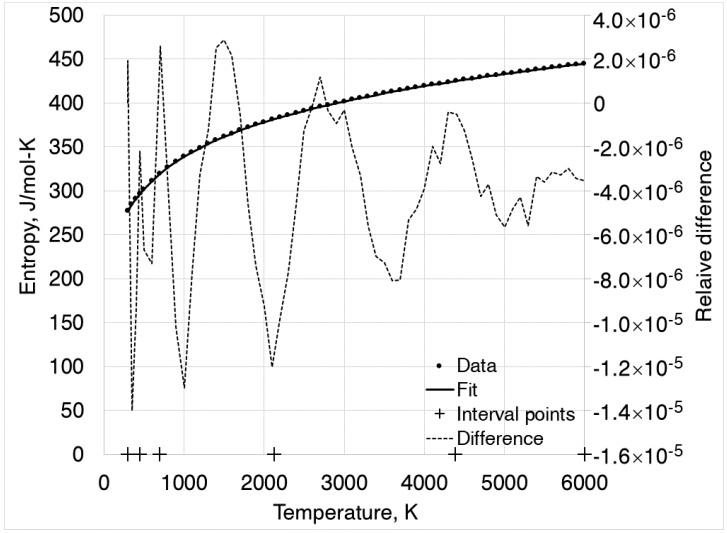


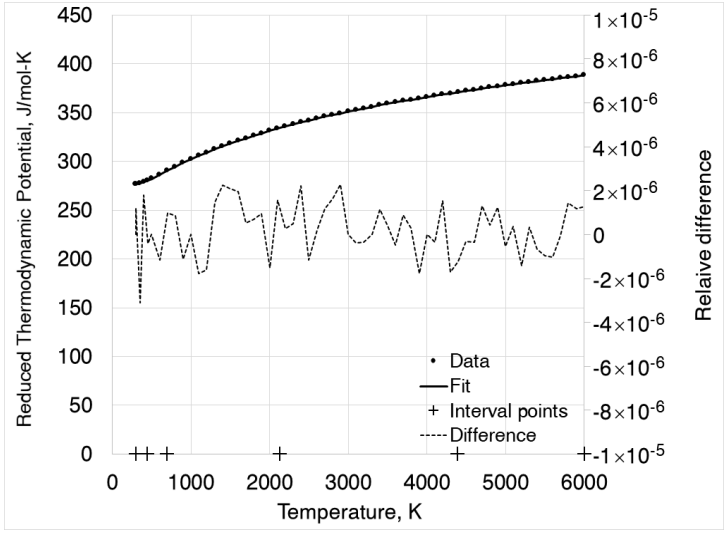
Molybdenum Oxide (MoO_2) – Mo1O2(g)

MW = 127.94 g/mol – $\Delta_f H^o$(298.15 K) = -8.314 kJ/mol – S^o(298.15 K) = 276.99 J/mol-K – p^o = 0.1 MPa



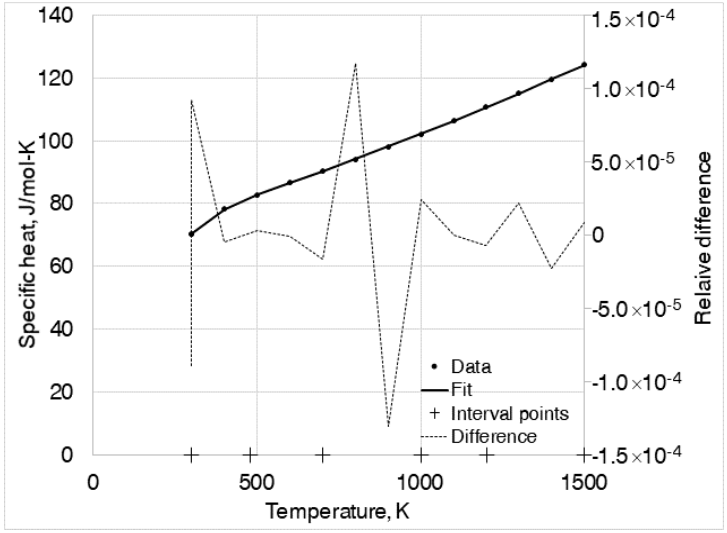


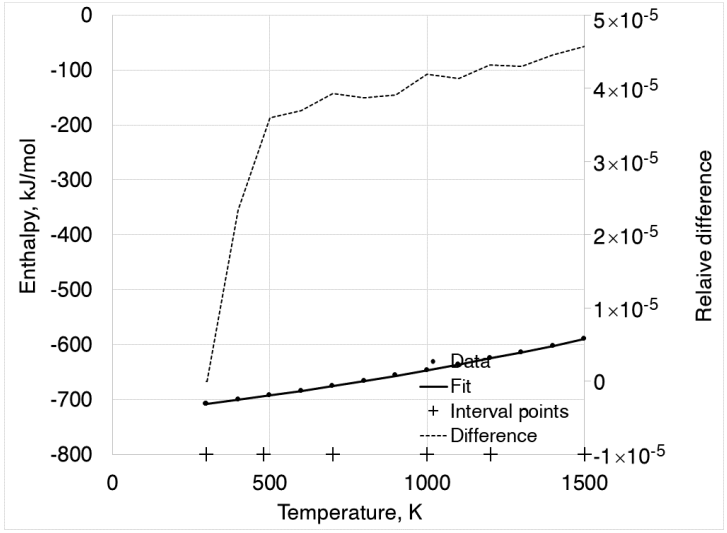


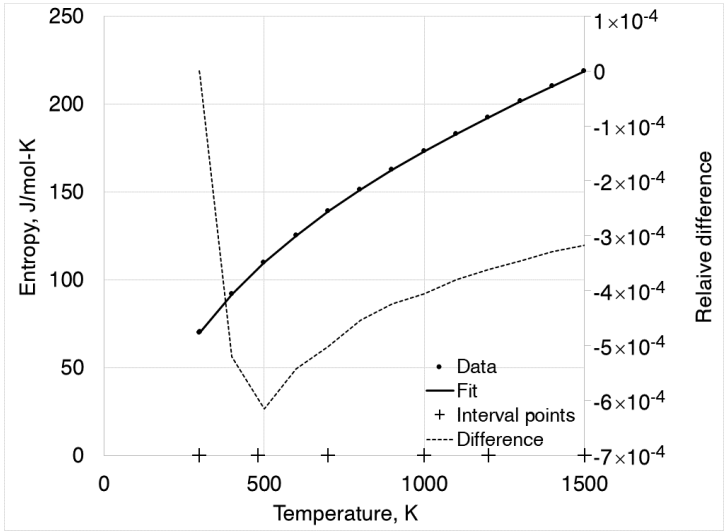


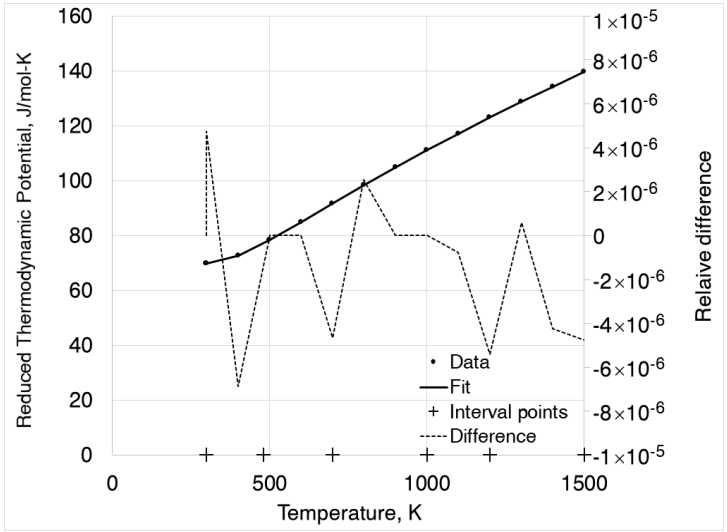
Molybdenum Oxide ($MoO_{2.750}$) – Mo1O2.750(cr)

MW = 139.94 g/mol – $\Delta_f H^o$(298.15 K) = -708.412 kJ/mol – S^o(298.15 K) = 69.936 J/mol-K – p^o = 0.1 MPa



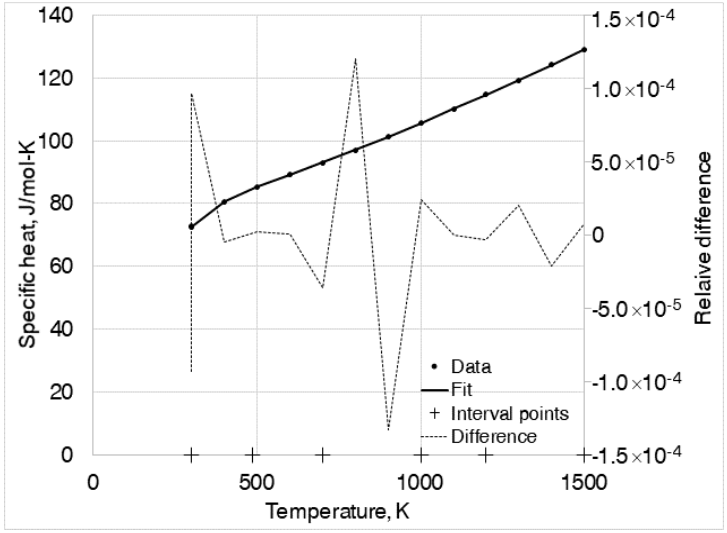


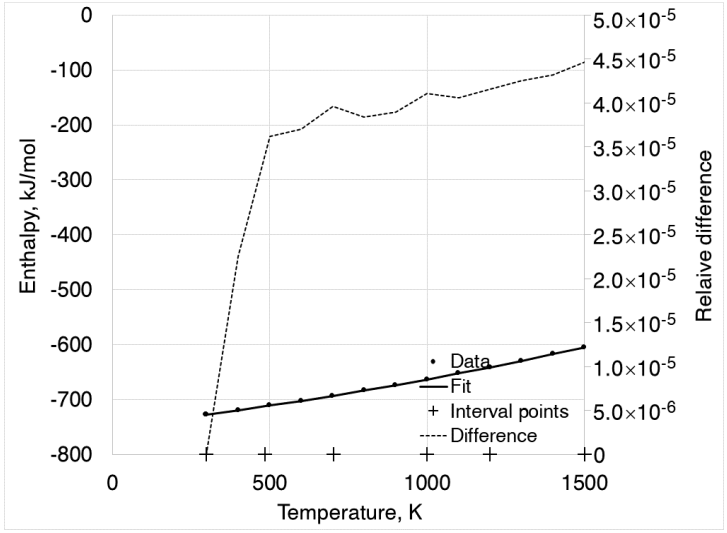


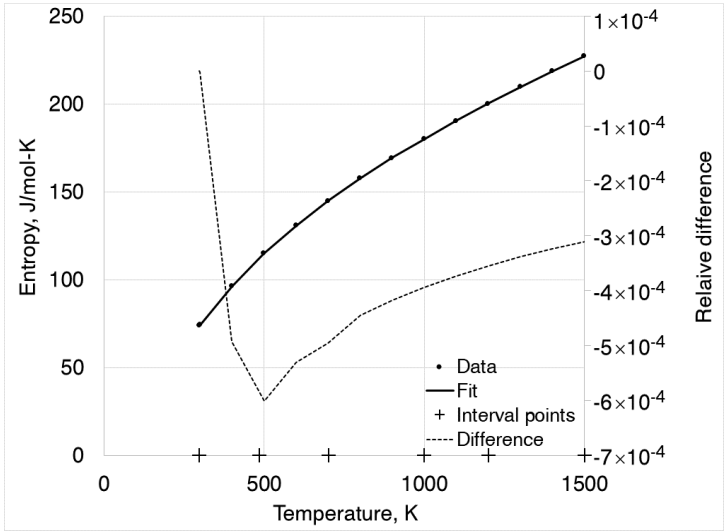


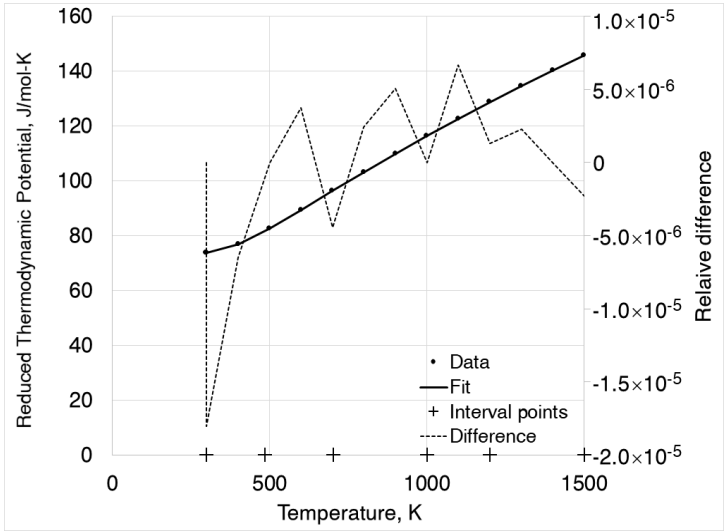
Molybdenum Oxide ($MoO_{2.875}$) – Mo1O2.875(cr)

MW = 141.94 g/mol – $\Delta_f H^o$(298.15 K) = -727.128 kJ/mol – S^o(298.15 K) = 73.849 J/mol-K – p^o = 0.1 MPa



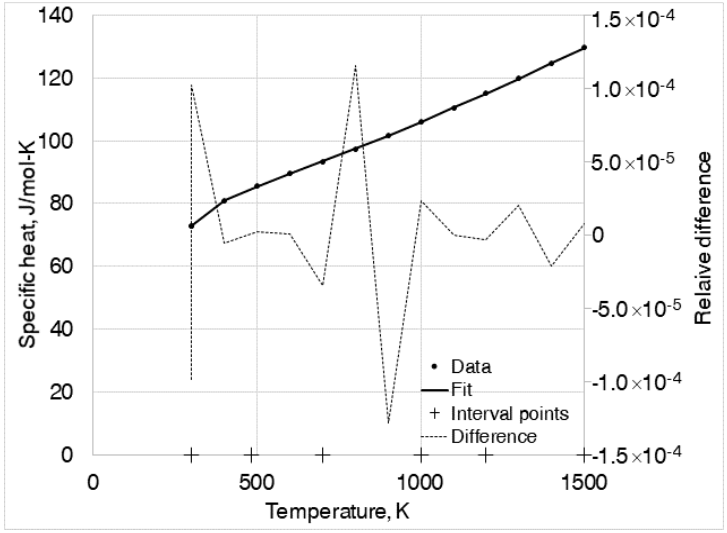


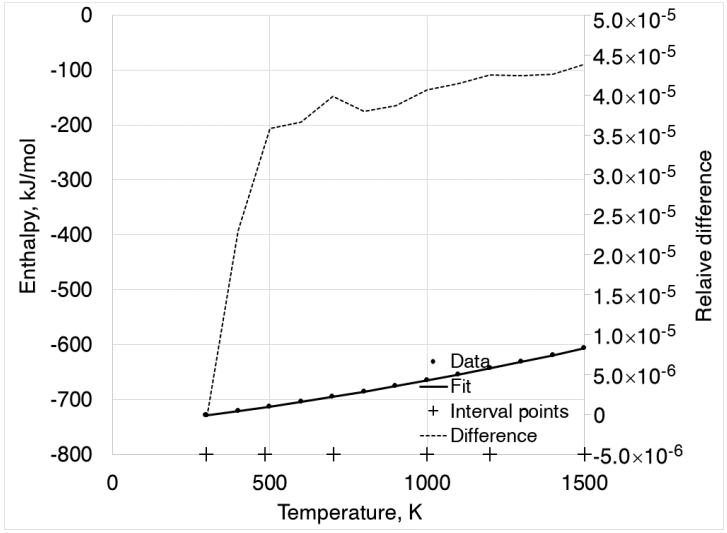


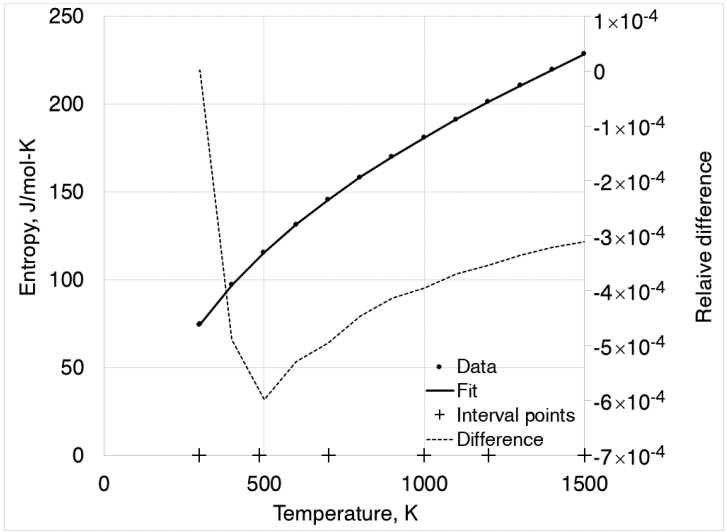


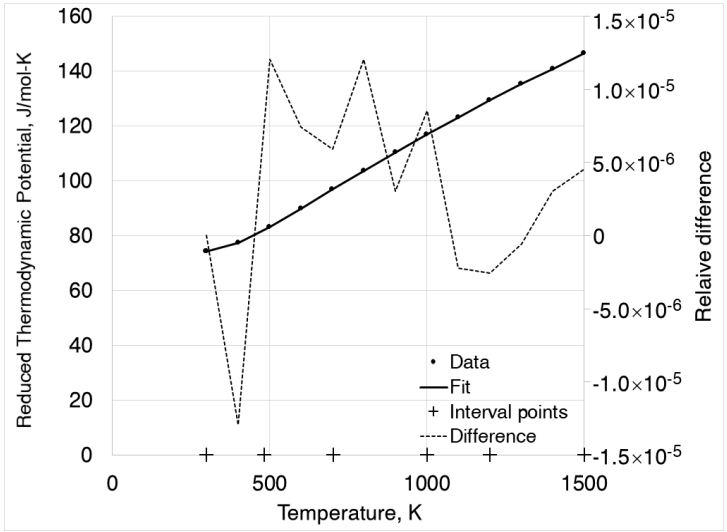
Molybdenum Oxide ($MoO_{2.889}$) – Mo1O2.889(cr)

MW = 142.16 g/mol – $\Delta_f H^o$(298.15 K) = -729.165 kJ/mol – S^o(298.15 K) = 74.281 J/mol-K – p^o = 0.1 MPa



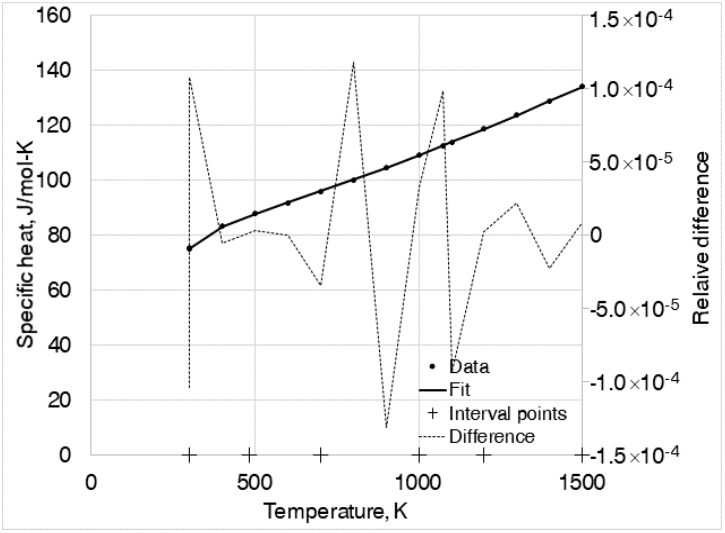


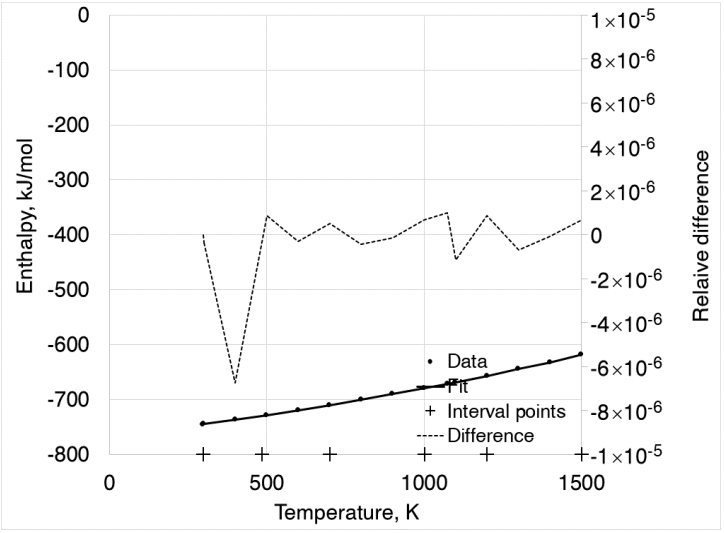


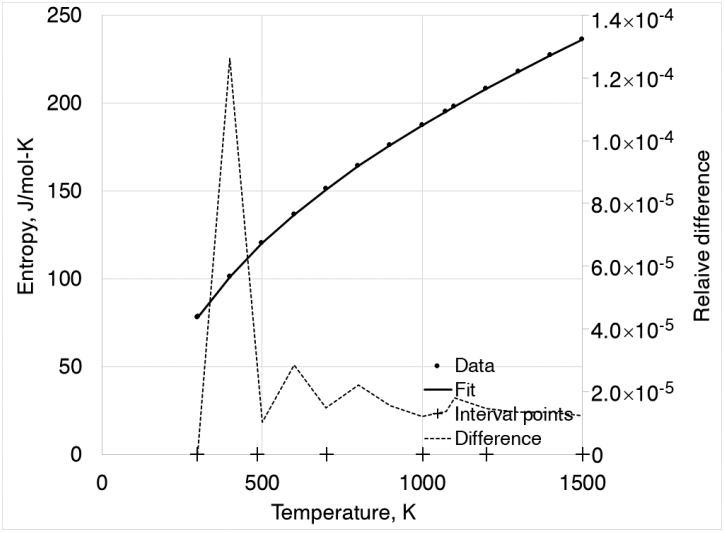


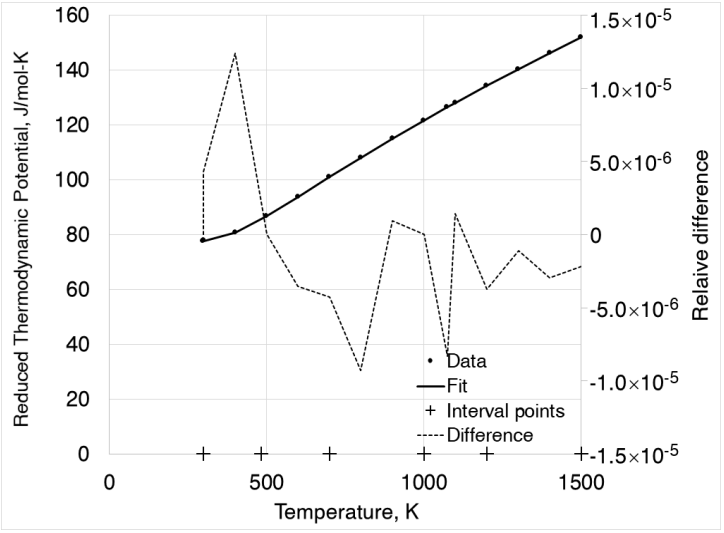
Molybdenum Oxide (MoO_3) – Mo1O3(cr)

MW = 143.94 g/mol – $\Delta_f H^o$(298.15 K) = -745.17 kJ/mol – S^o(298.15 K) = 77.76 J/mol-K – p^o = 0.1 MPa



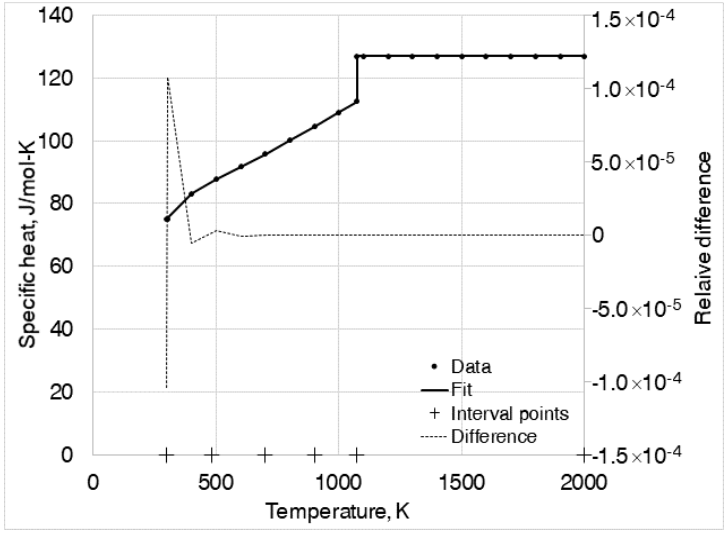


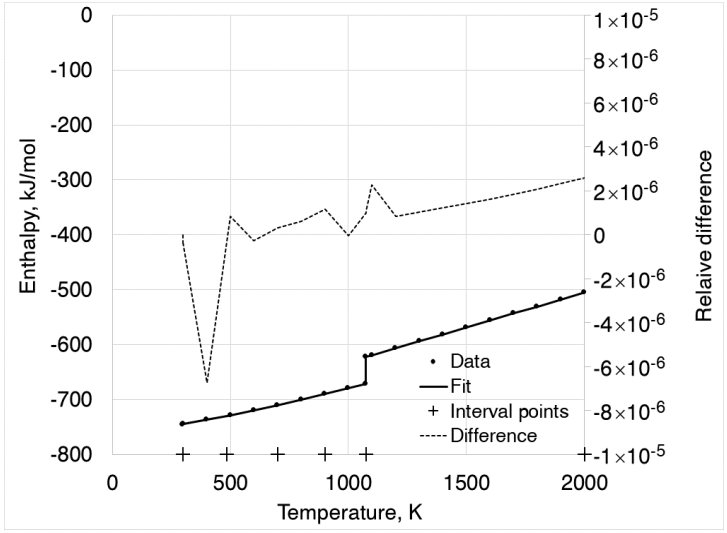


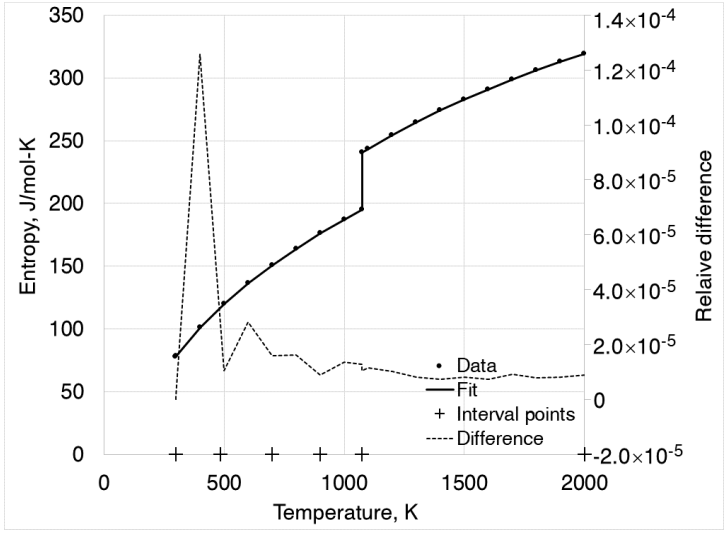


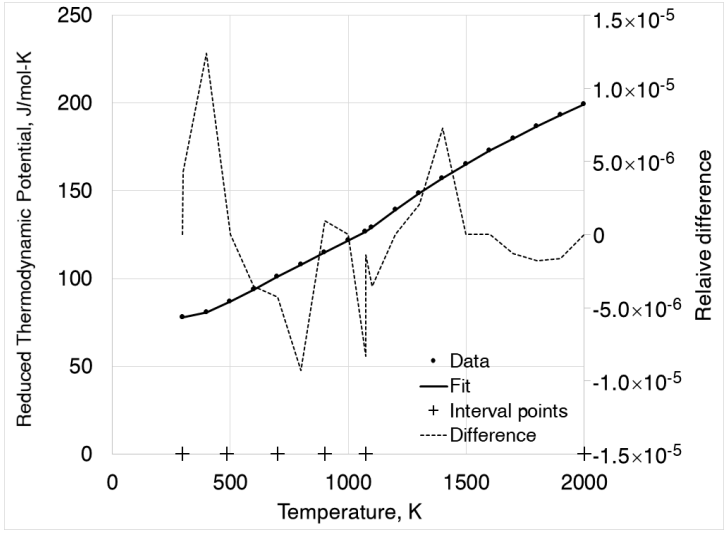
Molybdenum Oxide (MoO_3) – Mo1O3(cr,l)

MW = 143.94 g/mol – $\Delta_f H^o$(298.15 K) = -745.17 kJ/mol – S^o(298.15 K) = 77.76 J/mol-K – p^o = 0.1 MPa



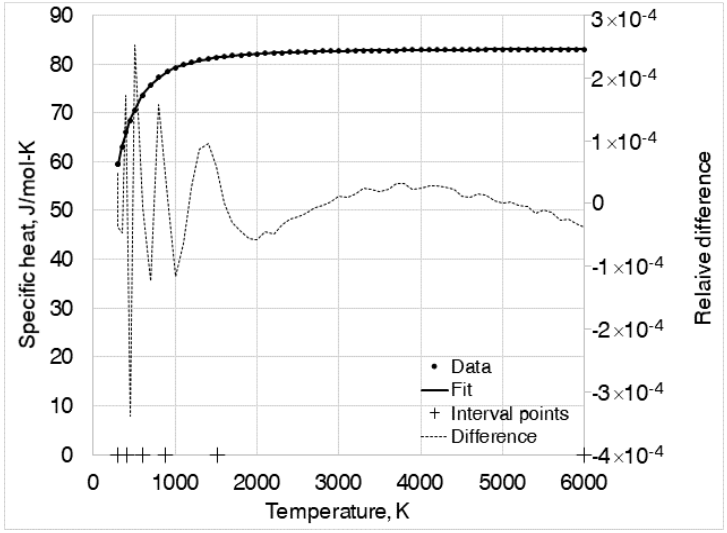


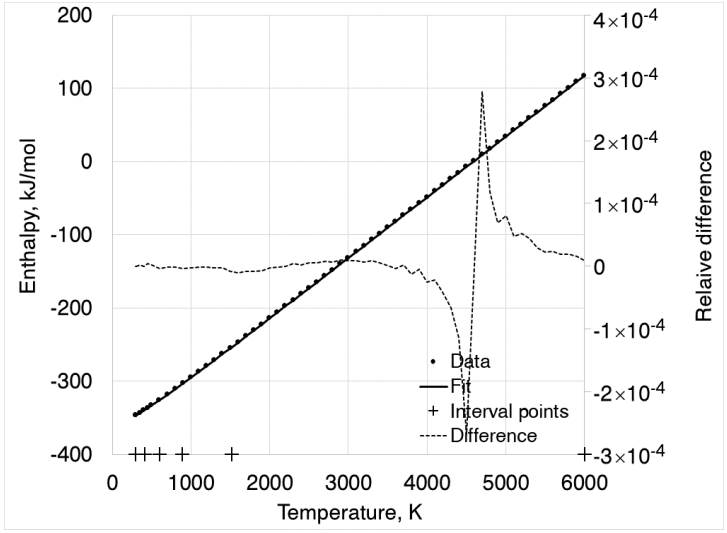


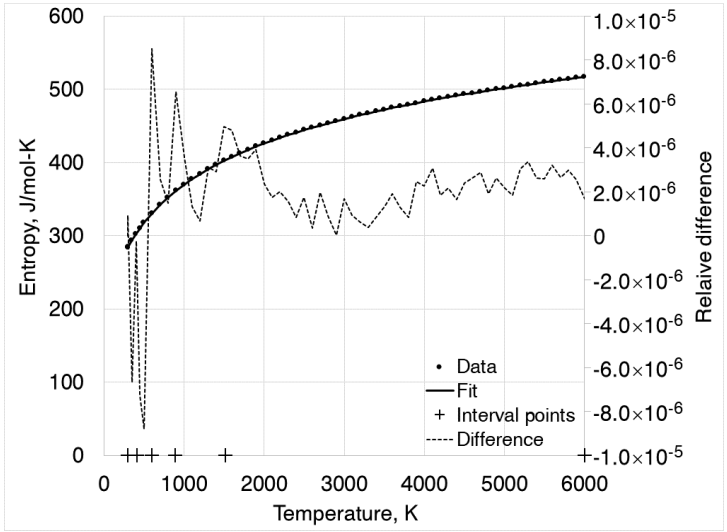


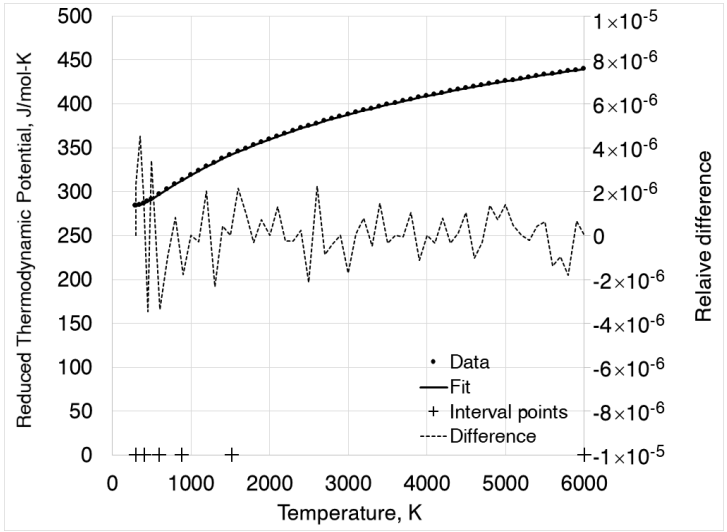
Molybdenum Oxide (MoO_3) – Mo1O3(g)

MW = 143.94 g/mol – $\Delta_f H^o$(298.15 K) = -346.435 kJ/mol – S^o(298.15 K) = 283.897 J/mol-K – p^o = 0.1 MPa



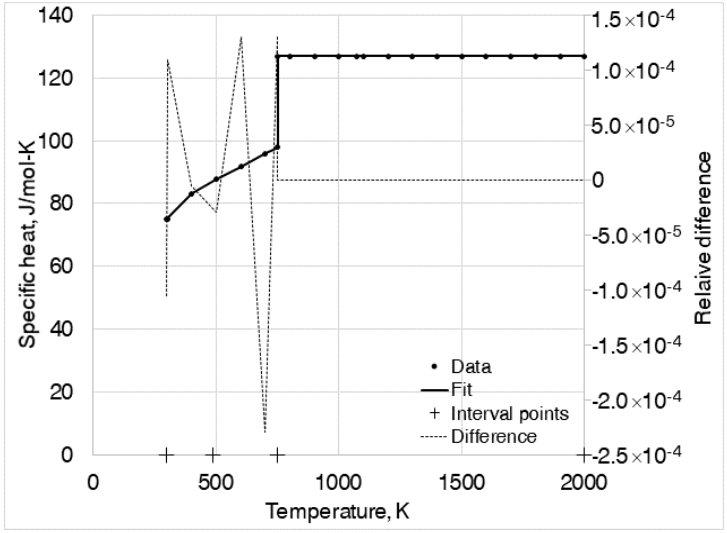


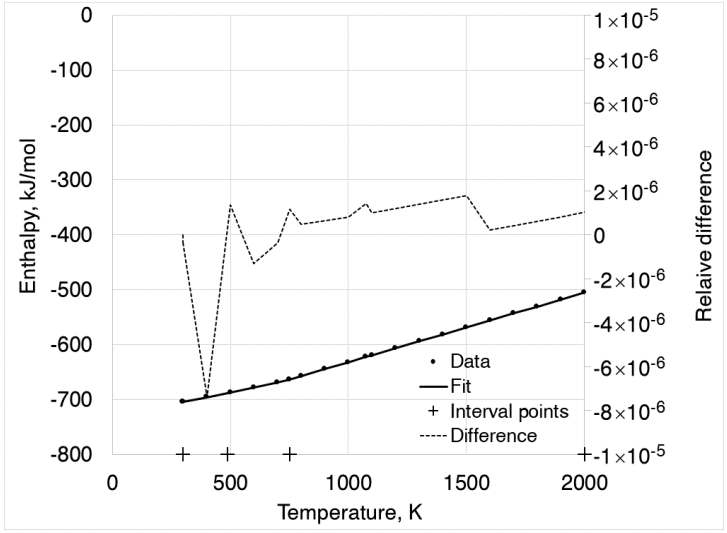


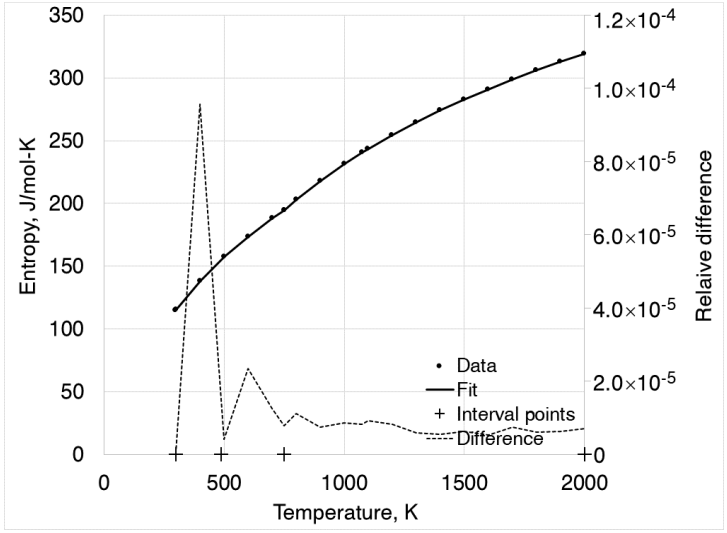


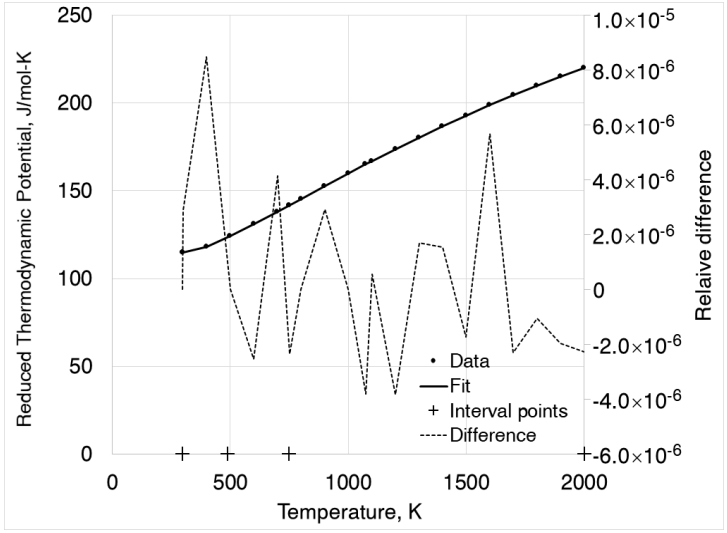
Molybdenum Oxide (MoO_3) – Mo1O3(l)

MW = 143.94 g/mol – $\Delta_f H^o$(298.15 K) = -703.727 kJ/mol – S^o(298.15 K) = 114.935 J/mol-K – p^o = 0.1 MPa



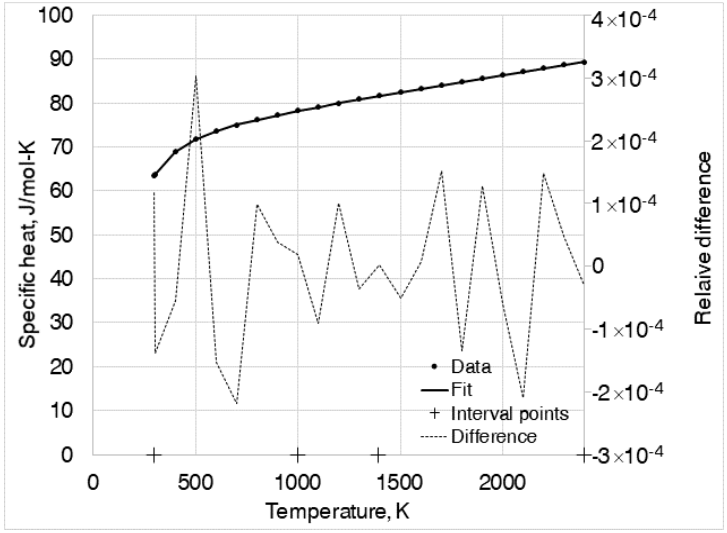


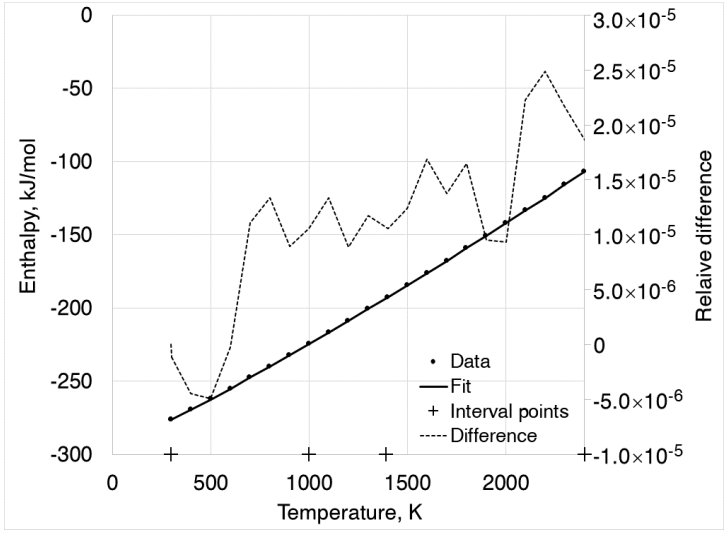


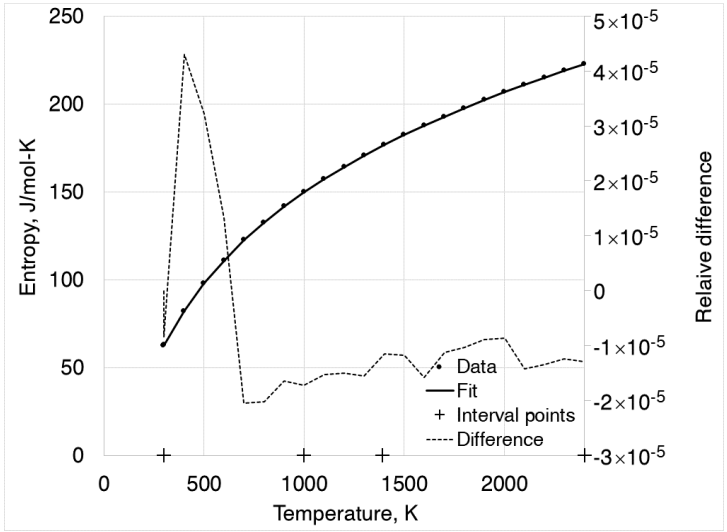


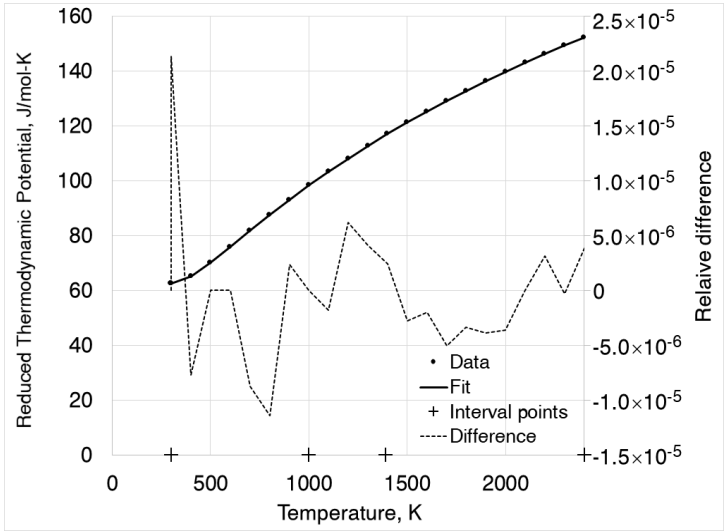
Molybdenum Sulfide (MoS_2) – Mo1S2(cr)

MW = 160.06 g/mol – $\Delta_f H^o$(298.15 K) = -276.144 kJ/mol – S^o(298.15 K) = 62.572 J/mol-K – p^o = 0.1 MPa



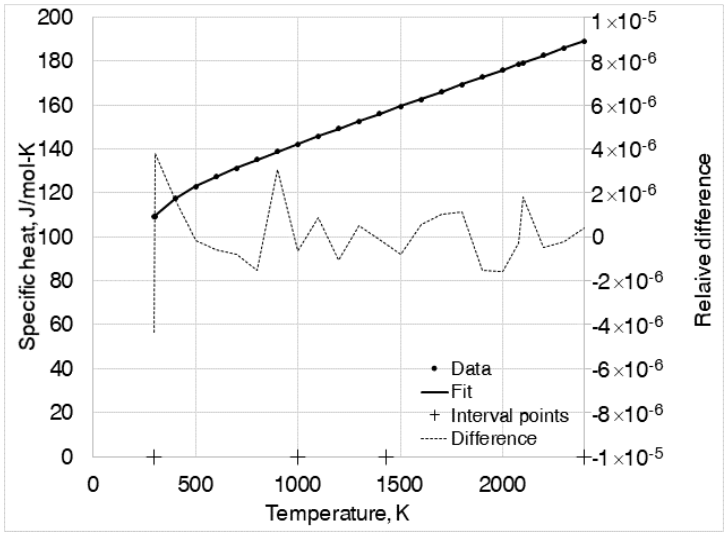


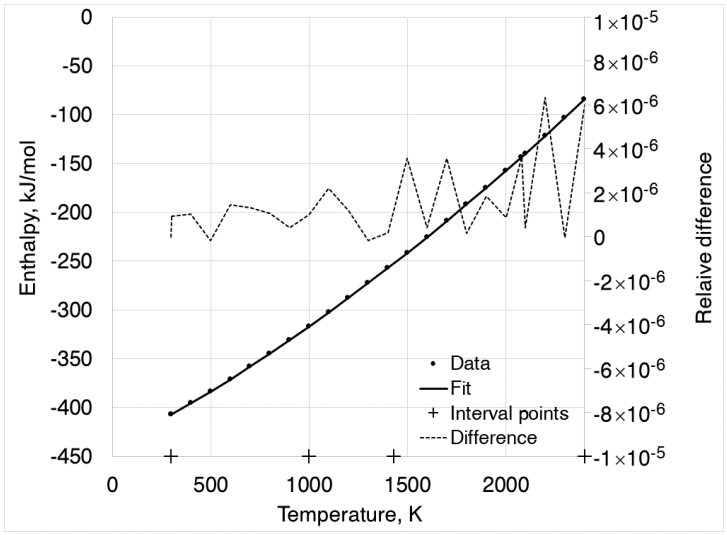


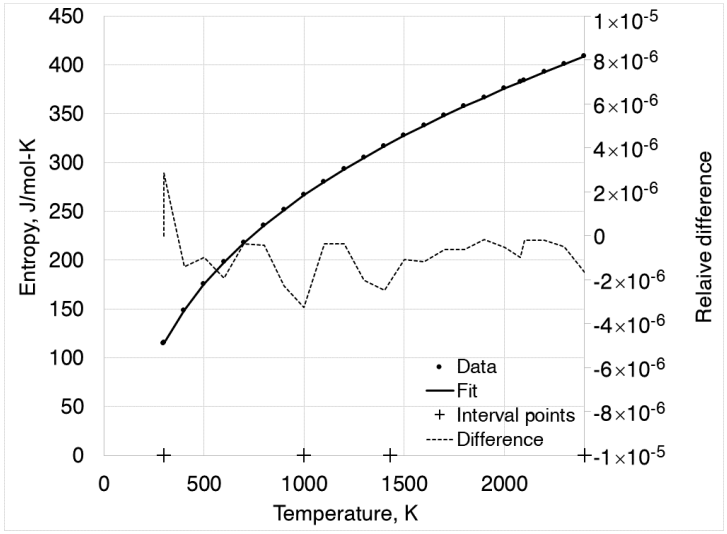


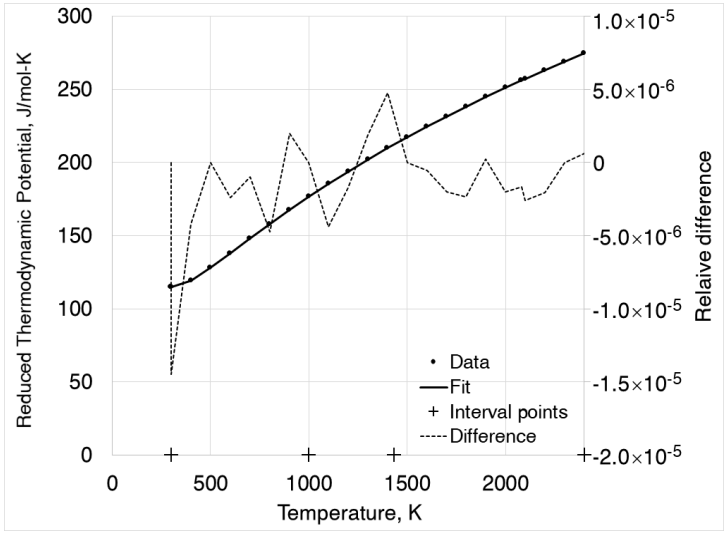
Molybdenum Sulfide (Mo_2S_3) – Mo2S3(cr)

MW = 288.06 g/mol – $\Delta_f H^o$(298.15 K) = -407.103 kJ/mol – S^o(298.15 K) = 114.976 J/mol-K – p^o = 0.1 MPa



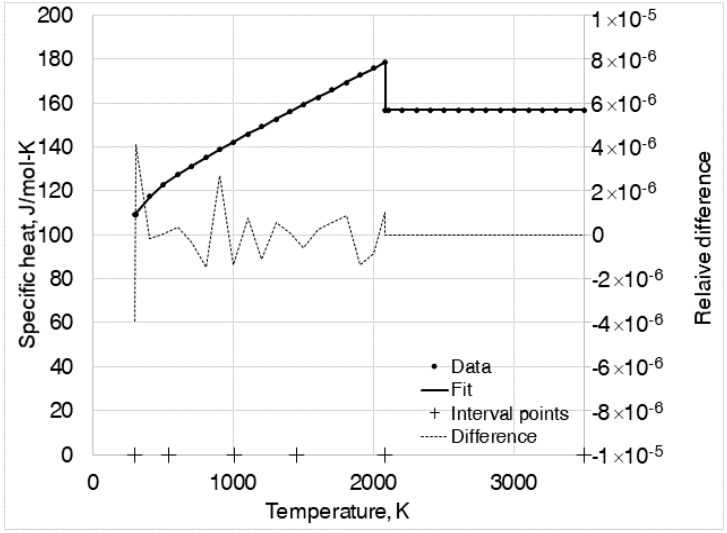


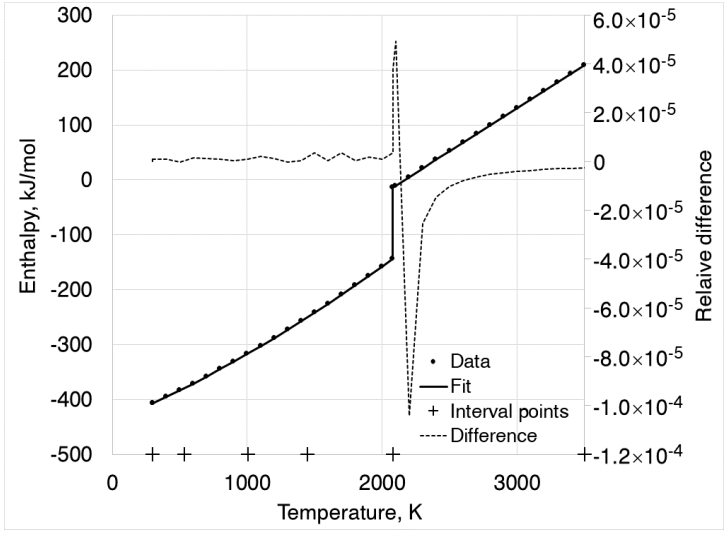


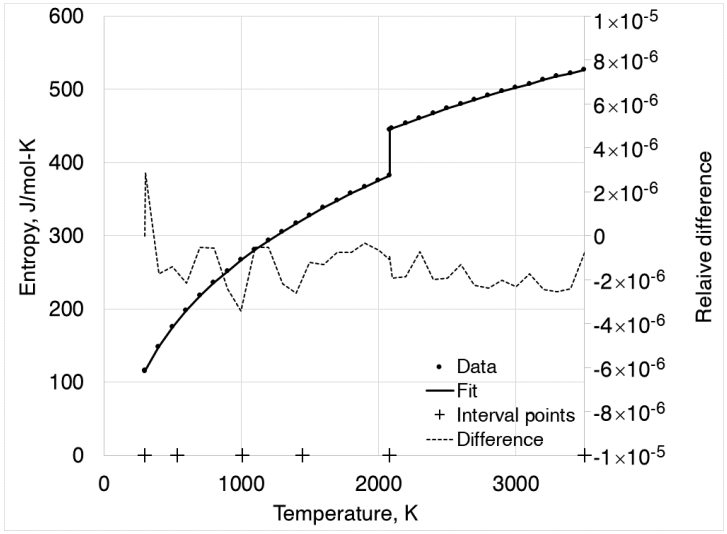


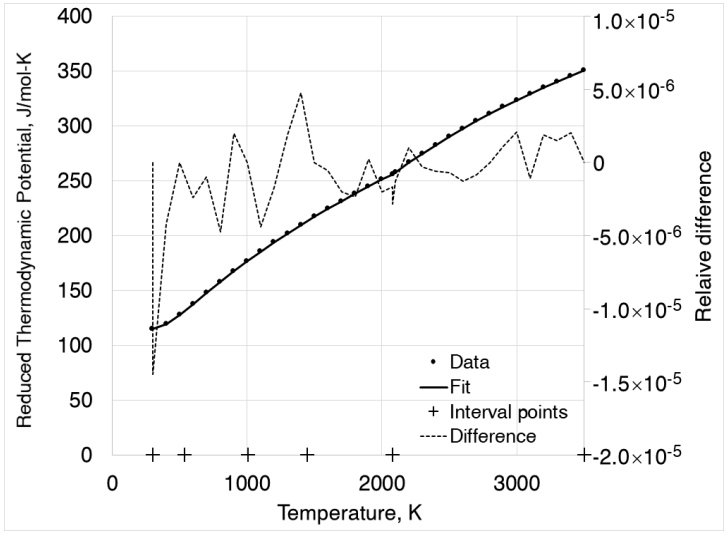
Molybdenum Sulfide (Mo_2S_3) – Mo2S3(cr,l)

MW = 288.06 g/mol – $\Delta_f H^o$(298.15 K) = -407.103 kJ/mol – S^o(298.15 K) = 114.976 J/mol-K – p^o = 0.1 MPa



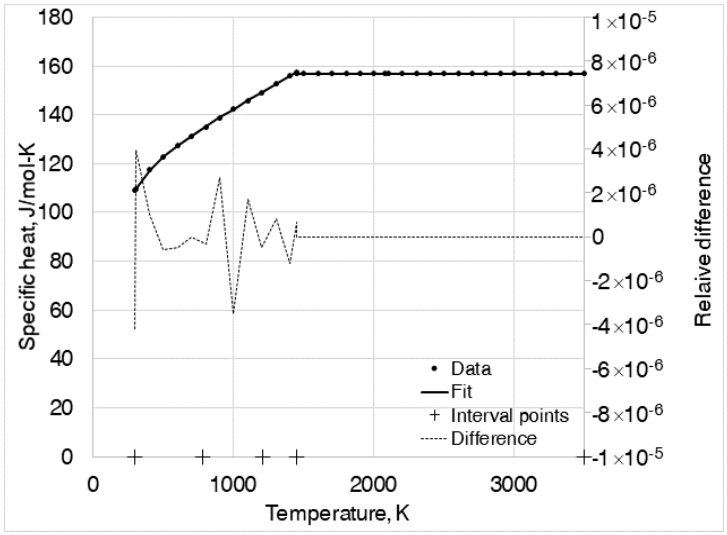


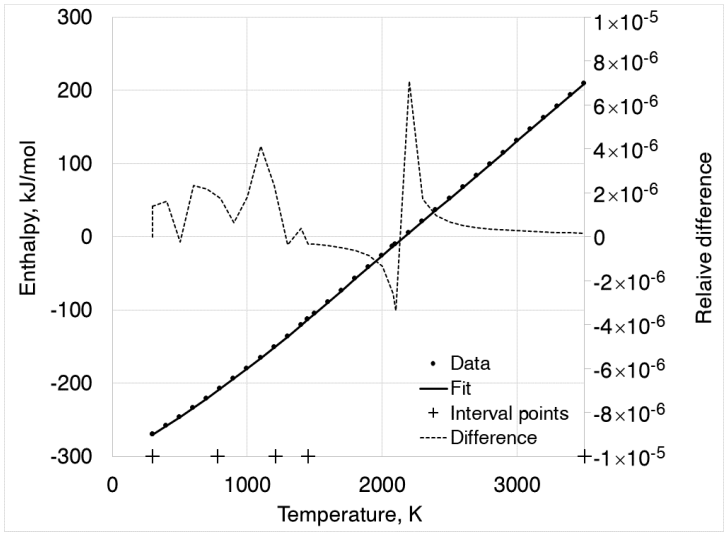


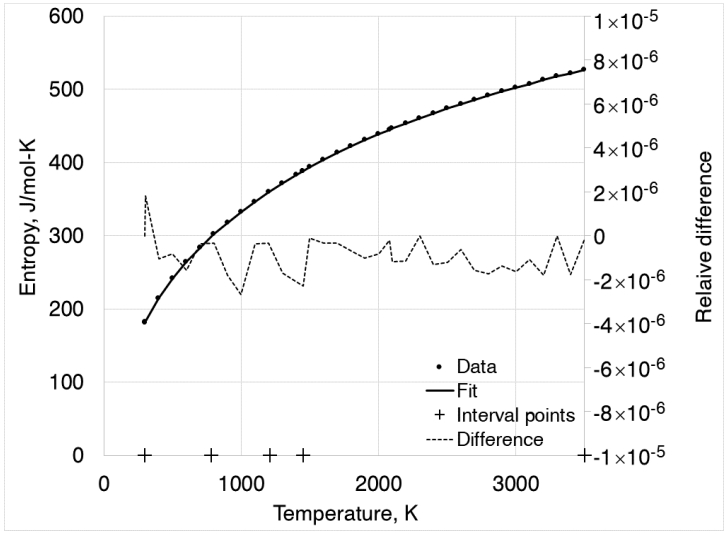


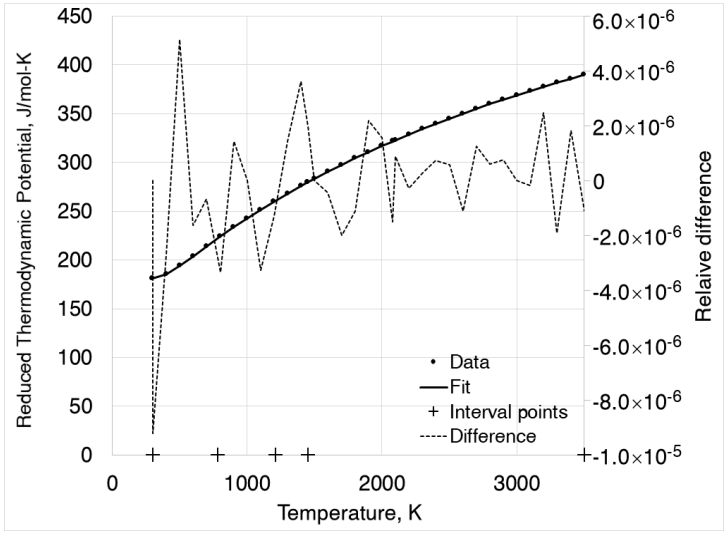
Molybdenum Sulfide (Mo_2S_3) – Mo2S3(l)

MW = 288.06 g/mol – $\Delta_f H^o$(298.15 K) = -270.304 kJ/mol – S^o(298.15 K) = 181.17 J/mol-K – p^o = 0.1 MPa



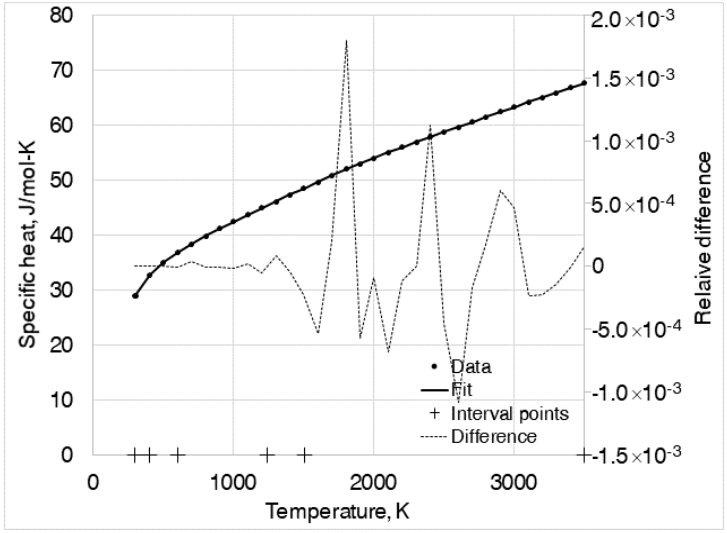


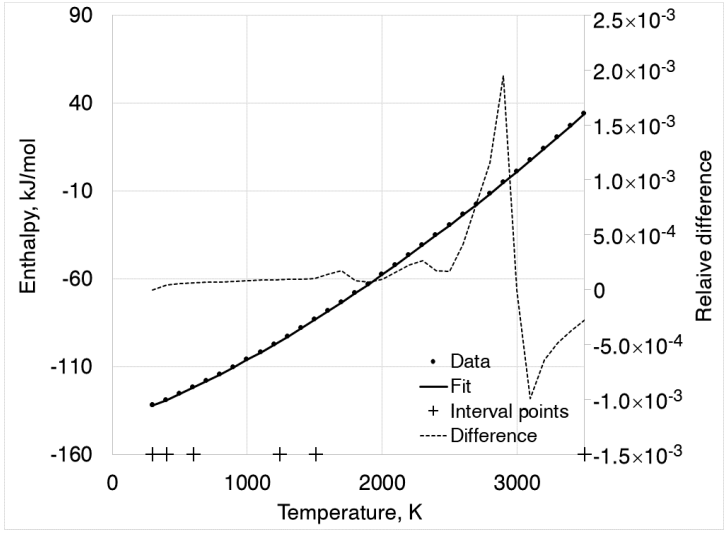


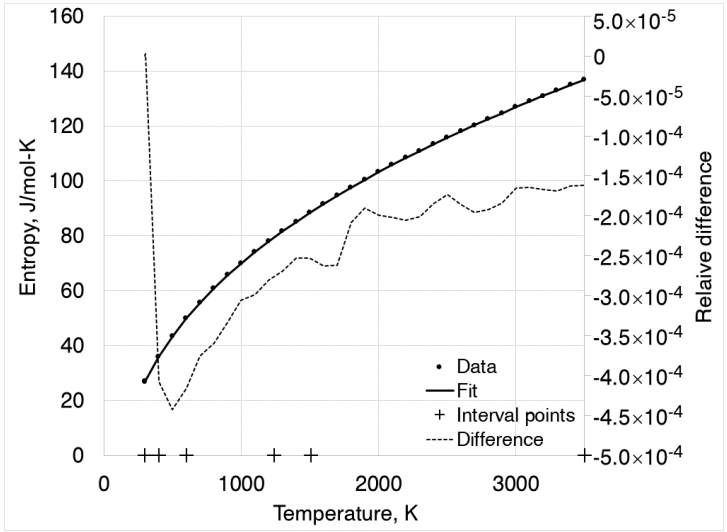


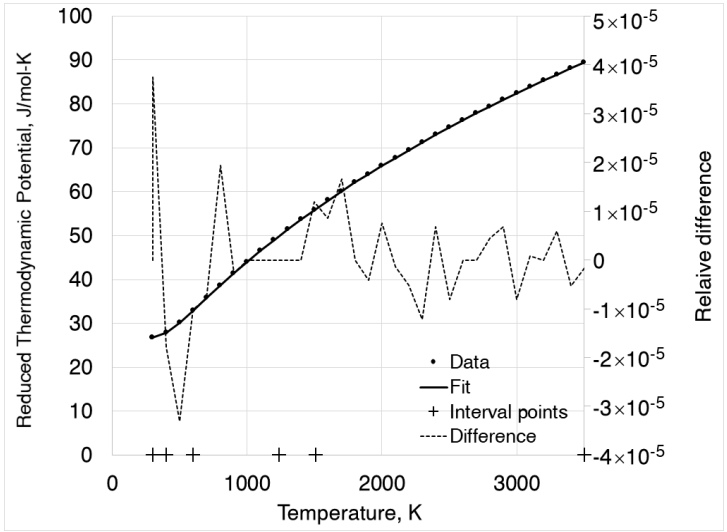
Vanadium Nitride ($VN_{0.465}$) – N0.465V1(cr)

MW = 57.4546 g/mol – $\Delta_f H^o$(298.15 K) = -132.214 kJ/mol – S^o(298.15 K) = 26.712 J/mol-K – p^o = 0.1 MPa



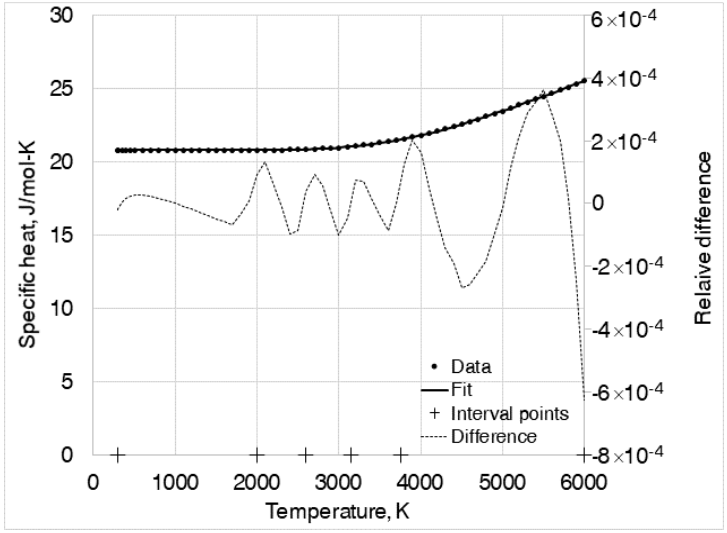


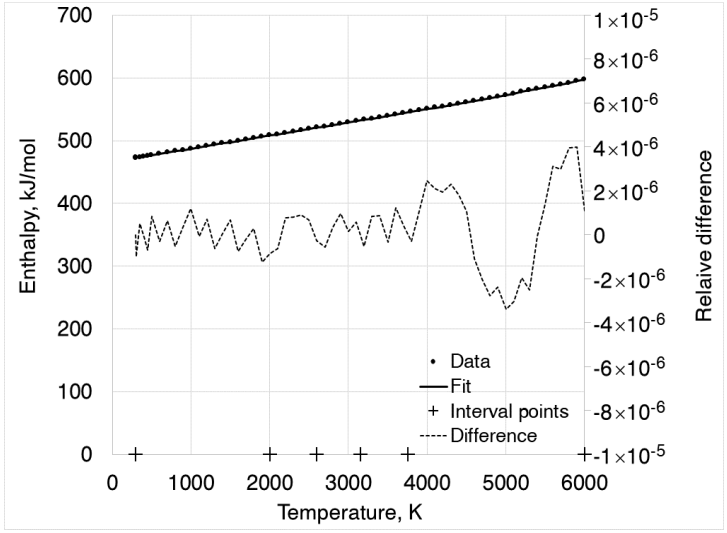


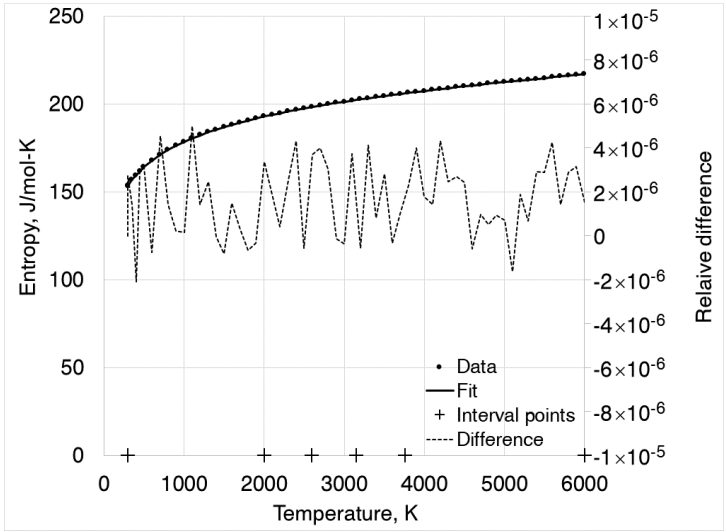


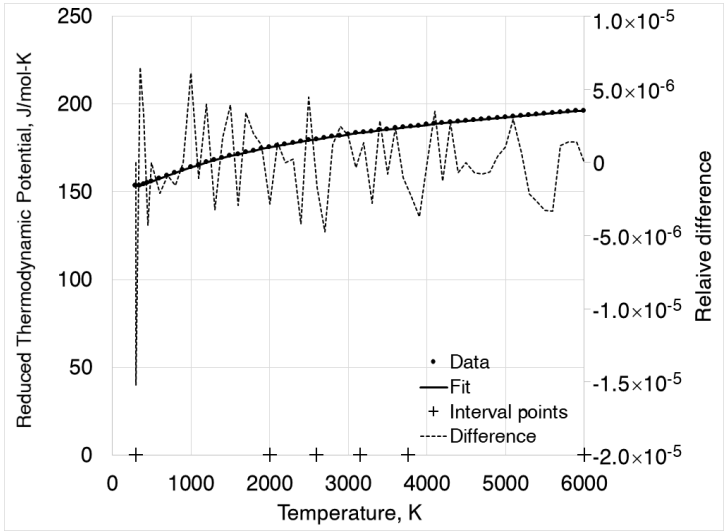
Nitrogen (N) – N1(g)

MW = 14.0067 g/mol – $\Delta_f H^o$(298.15 K) = 472.683 kJ/mol – S^o(298.15 K) = 153.3 J/mol-K – p^o = 0.1 MPa



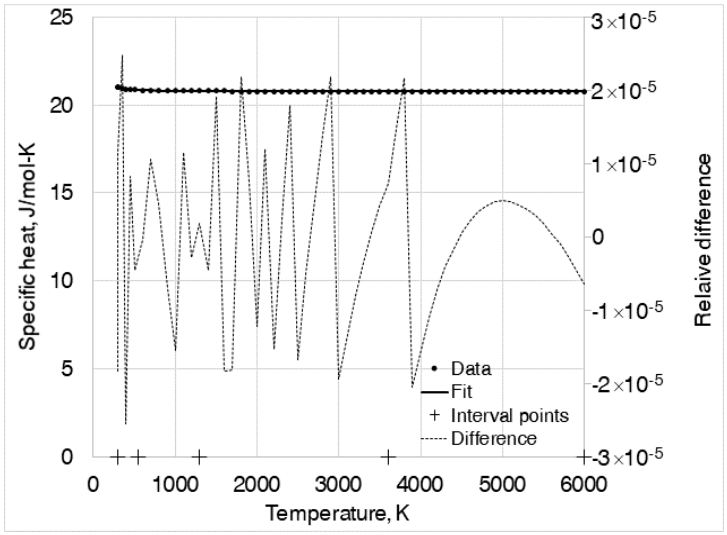


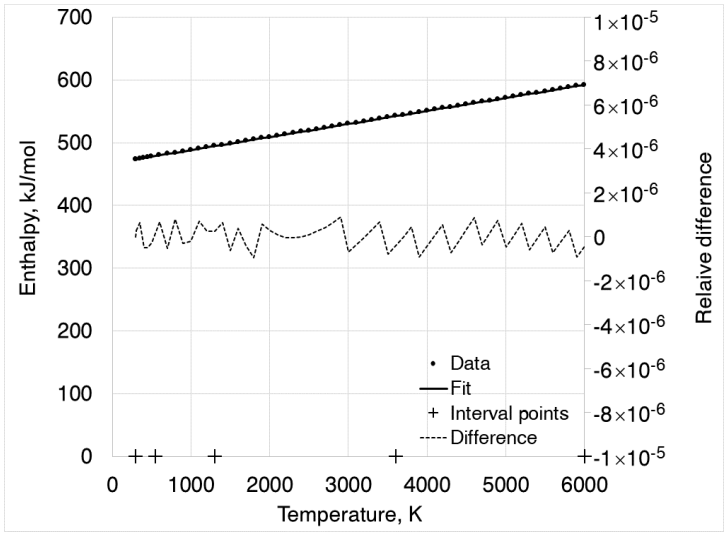


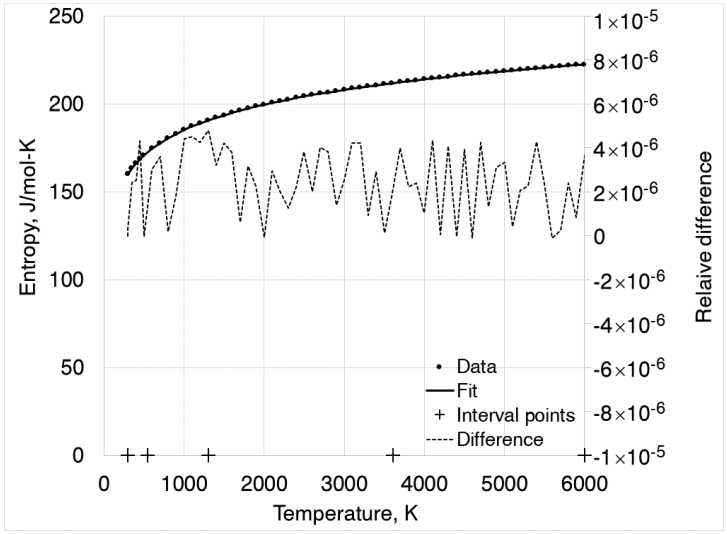


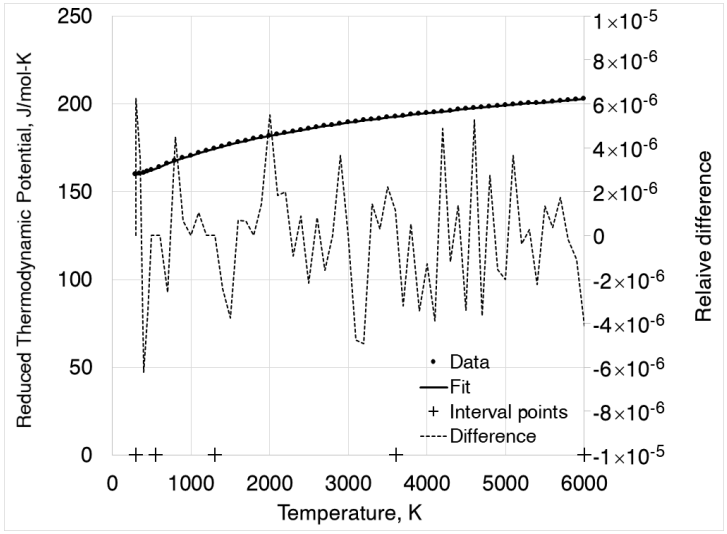
Nitrogen, Ion (N^-) – N1-(g)

MW = 14.0072 g/mol – $\Delta_f H^o$(298.15 K) = 473.54 kJ/mol – S^o(298.15 K) = 159.927 J/mol-K – p^o = 0.1 MPa



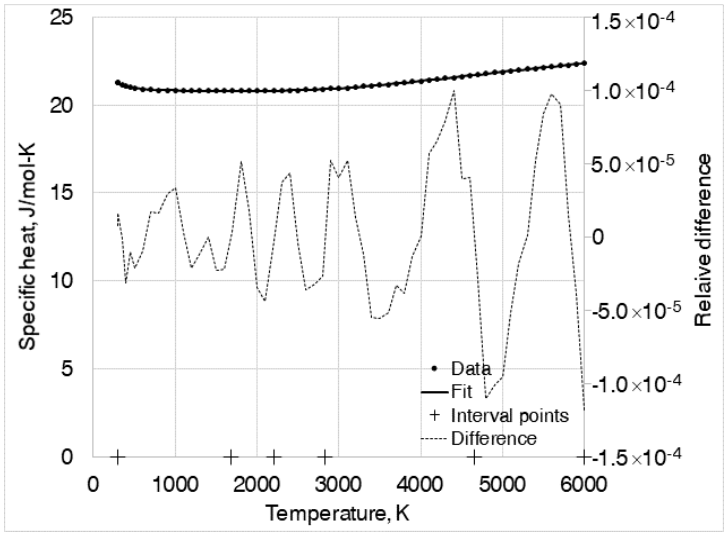


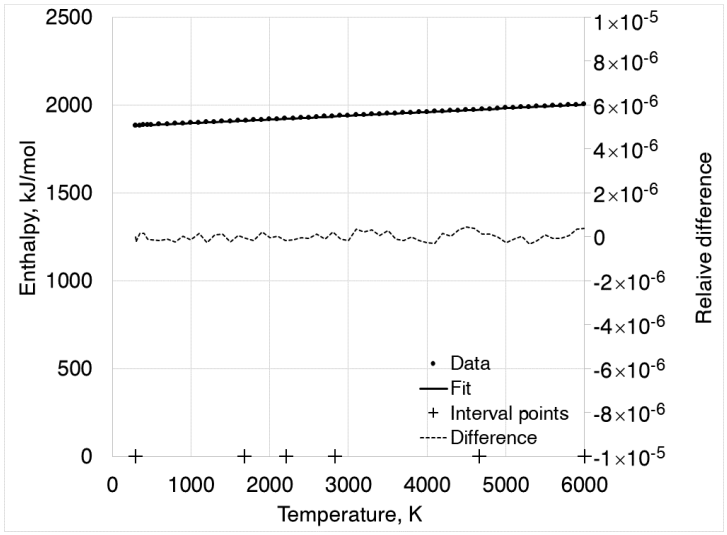


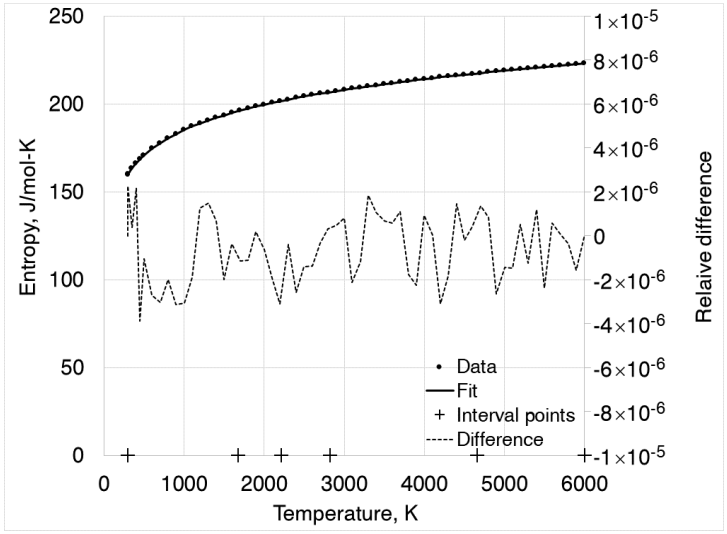


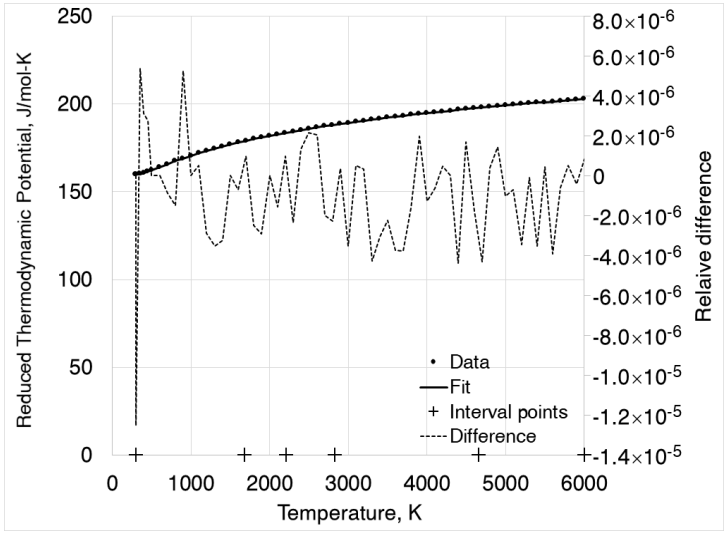
Nitrogen, Ion (N^+) – N1+(g)

MW = 14.0062 g/mol – $\Delta_f H^o$(298.15 K) = 1882.13 kJ/mol – S^o(298.15 K) = 159.797 J/mol-K – p^o = 0.1 MPa



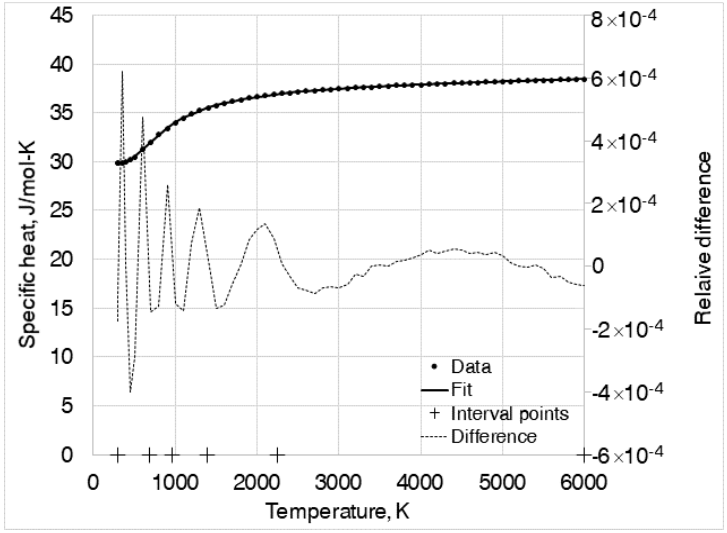


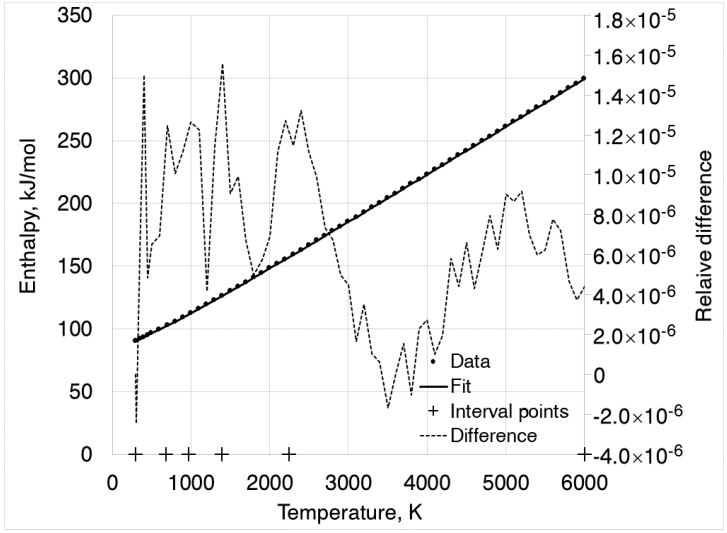


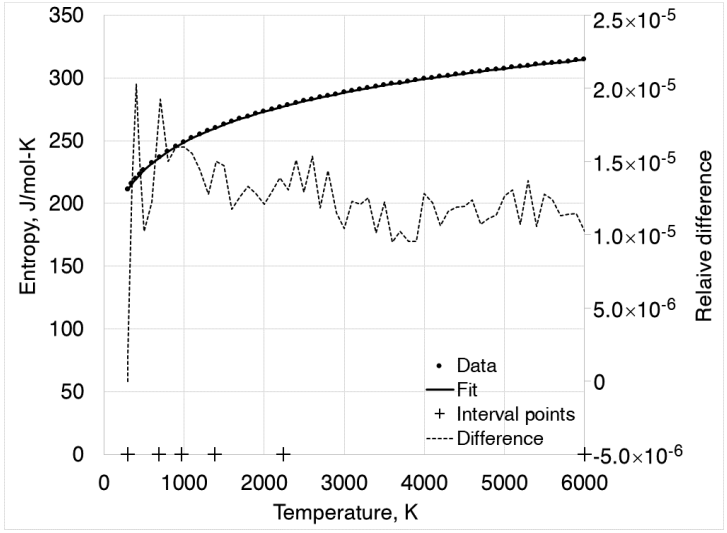


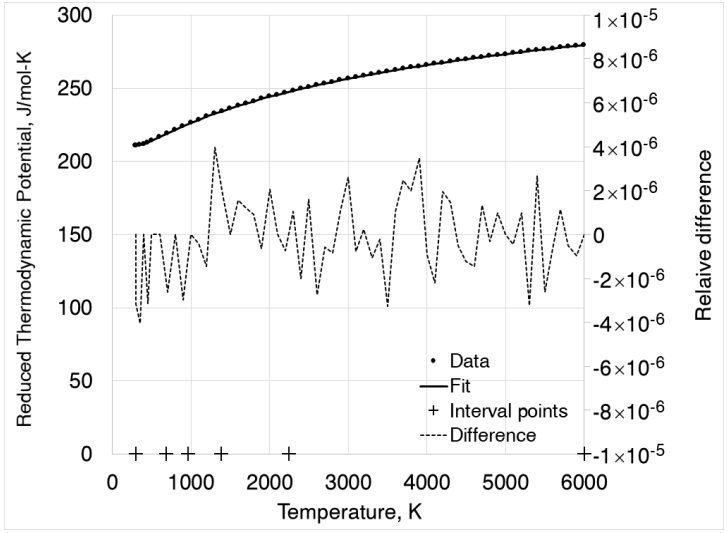
Nitrogen Oxide (NO) – N1O1(g)

MW = 30.0061 g/mol – $\Delta_f H^o$(298.15 K) = 90.291 kJ/mol – S^o(298.15 K) = 210.758 J/mol-K – p^o = 0.1 MPa



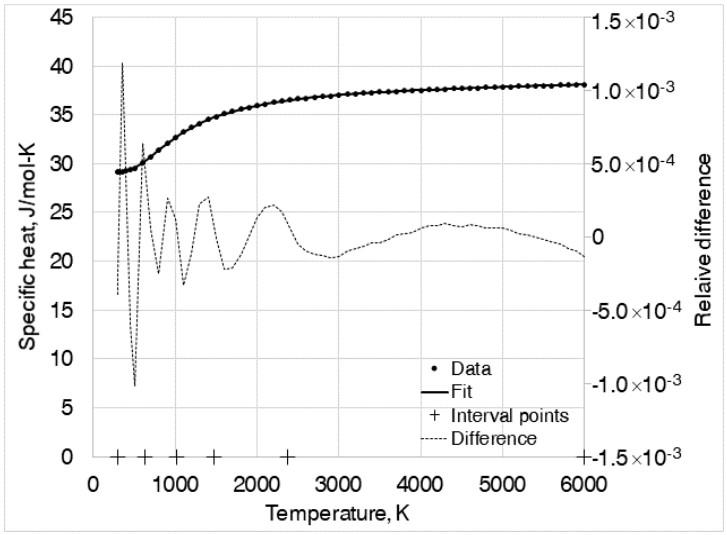


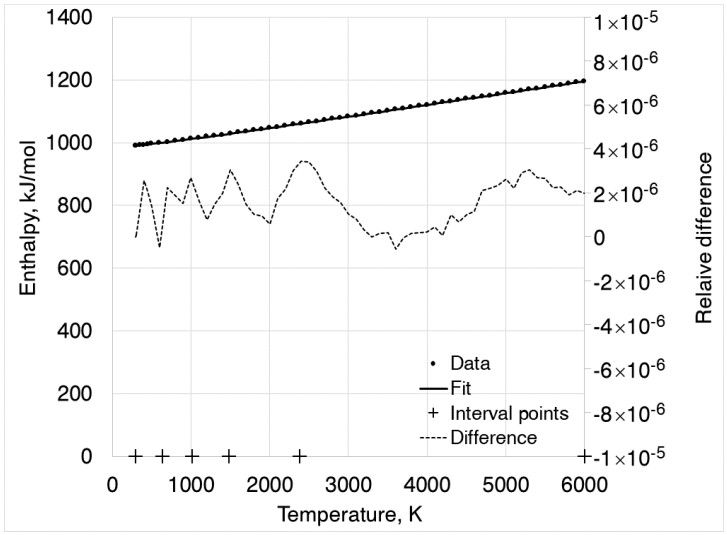


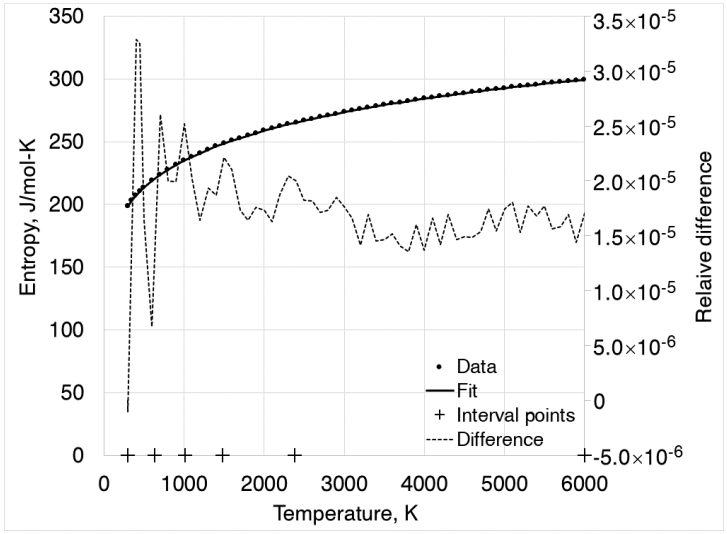


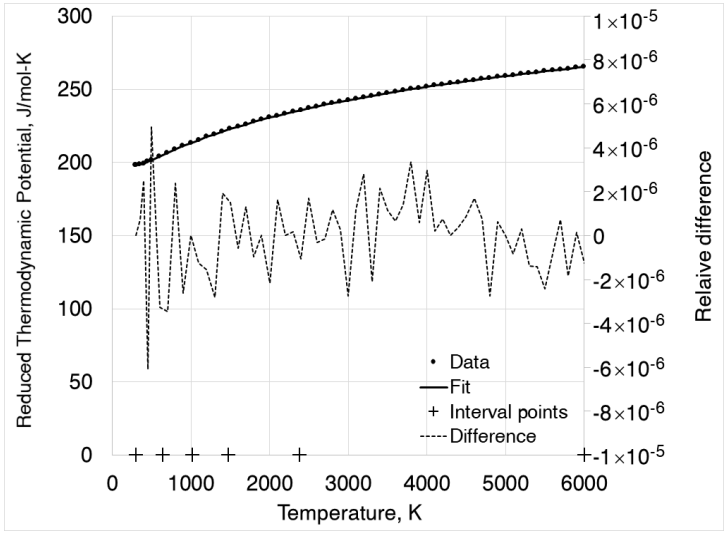
Nitrogen Oxide, Ion (NO^+) – N1O1+(g)

MW = 30.0056 g/mol – $\Delta_f H^o$(298.15 K) = 990.185 kJ/mol – S^o(298.15 K) = 198.219 J/mol-K – p^o = 0.1 MPa



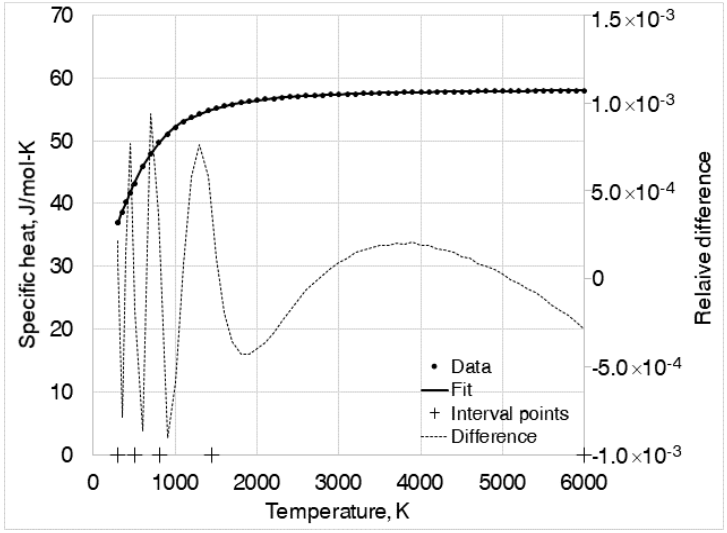


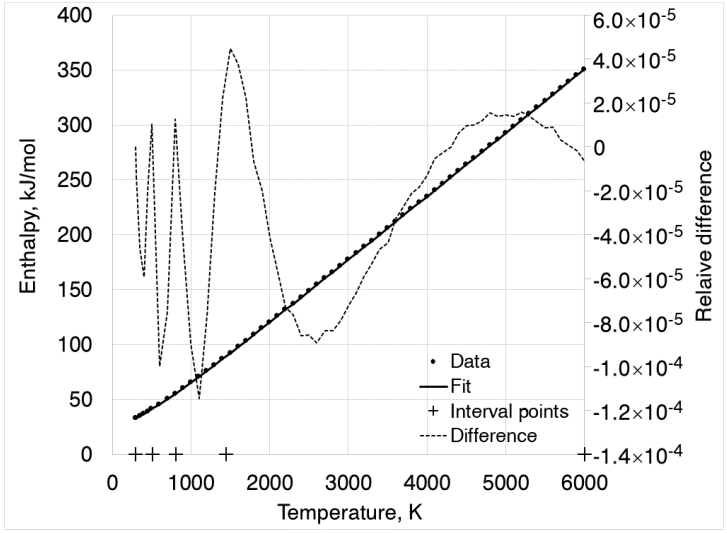


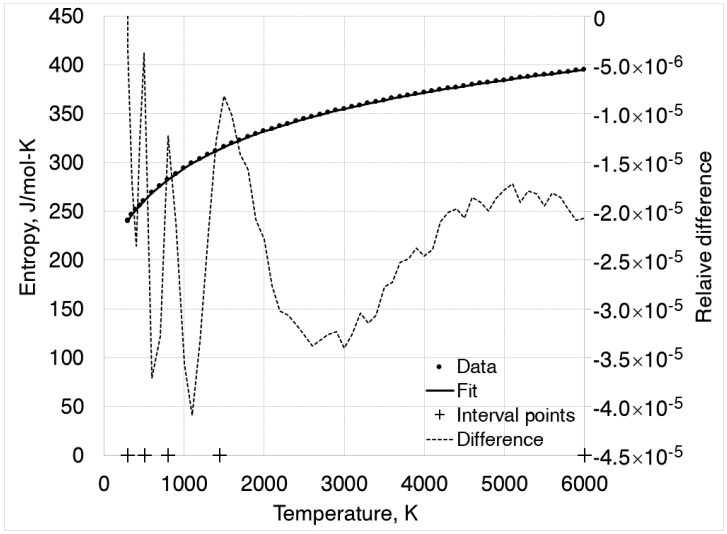


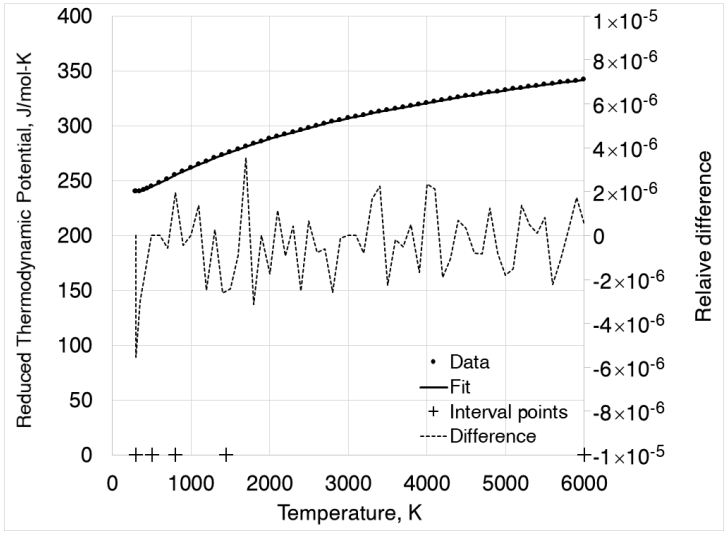
Nitrogen Oxide (NO_2) – N1O2(g)

MW = 46.0055 g/mol – $\Delta_f H^o$(298.15 K) = 33.095 kJ/mol – S^o(298.15 K) = 240.034 J/mol-K – p^o = 0.1 MPa



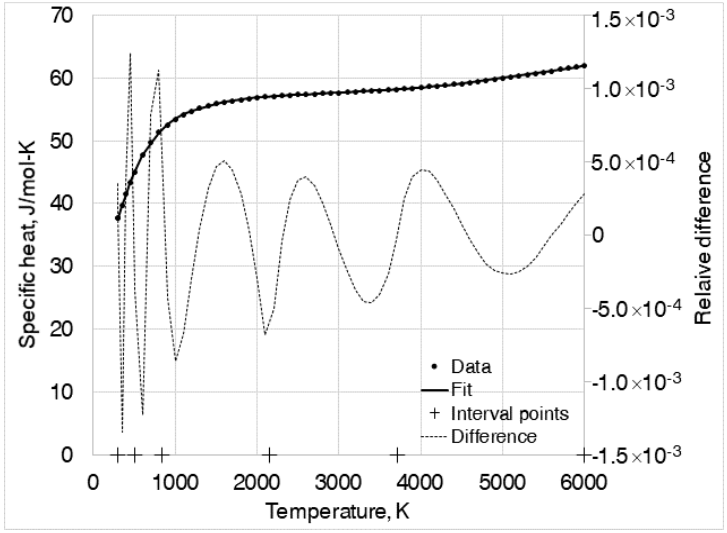


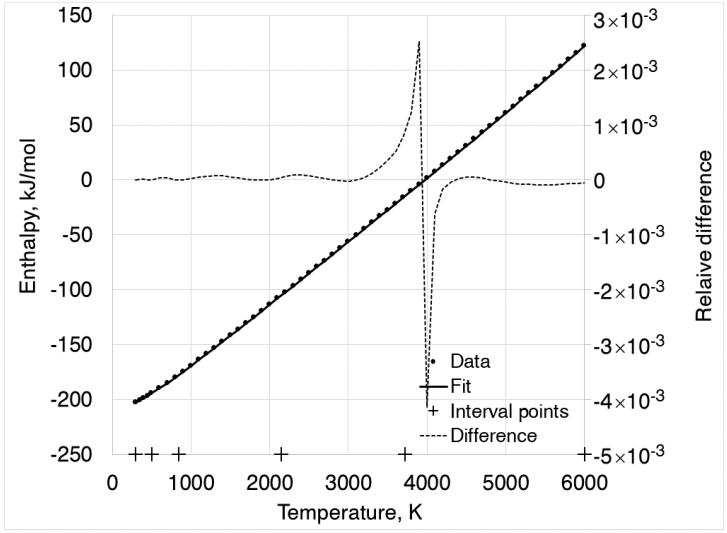


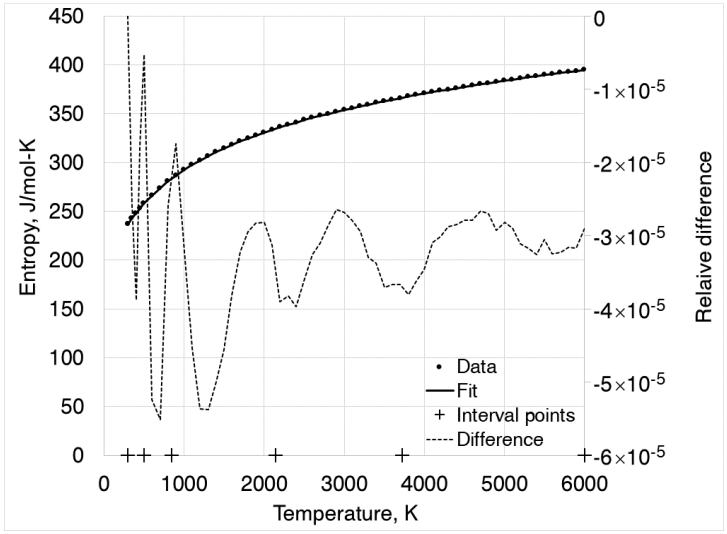


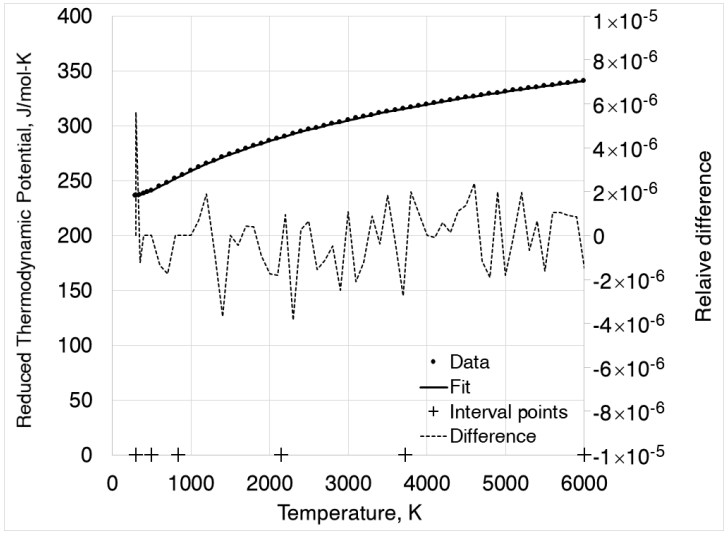
Nitrogen Oxide, Ion (ONO^-) – N1O2-(g)

MW = 46.006 g/mol – $\Delta_f H^o$(298.15 K) = -202.715 kJ/mol – S^o(298.15 K) = 236.598 J/mol-K – p^o = 0.1 MPa



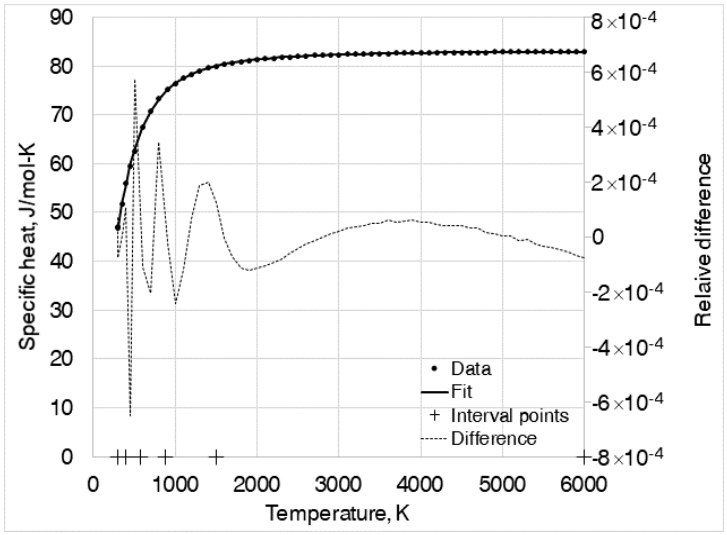


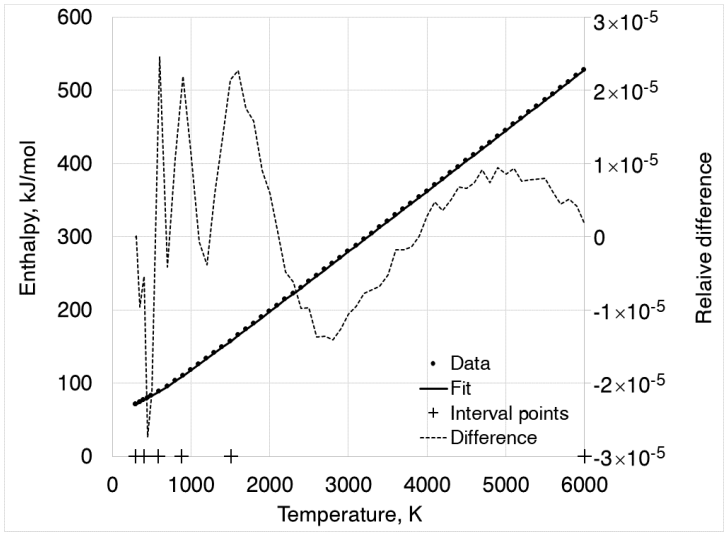


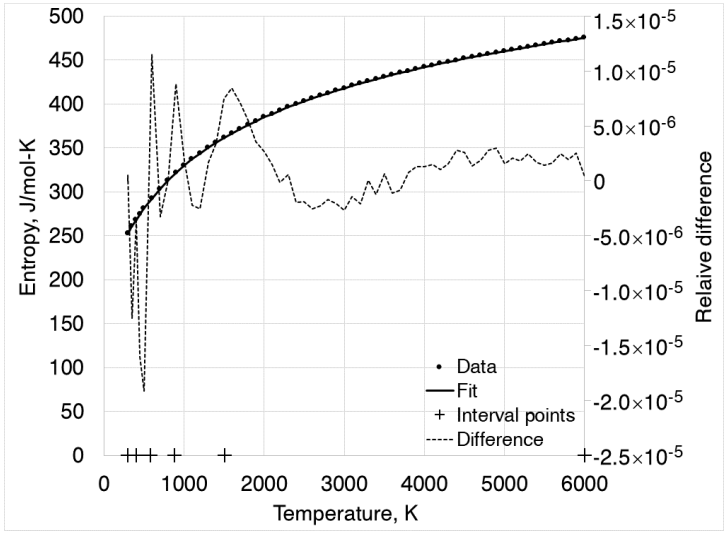


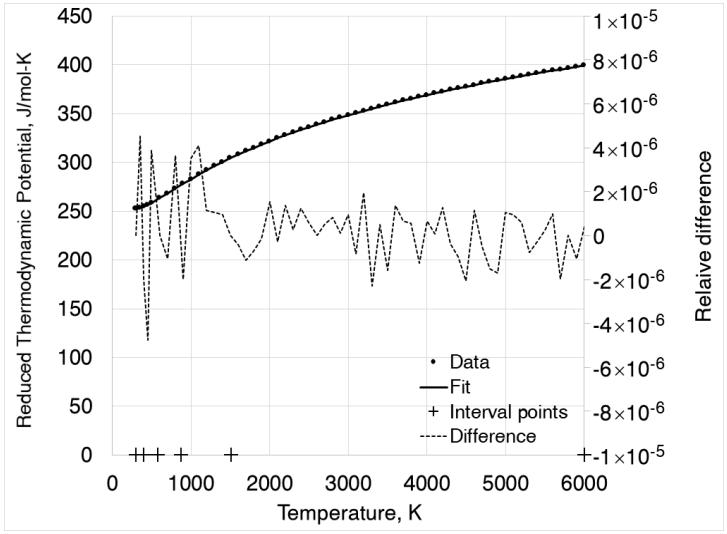
Nitrogen Oxide (NO_3) – N1O3(g)

MW = 62.0049 g/mol – $\Delta_f H^o$(298.15 K) = 71.128 kJ/mol – S^o(298.15 K) = 252.619 J/mol-K – p^o = 0.1 MPa



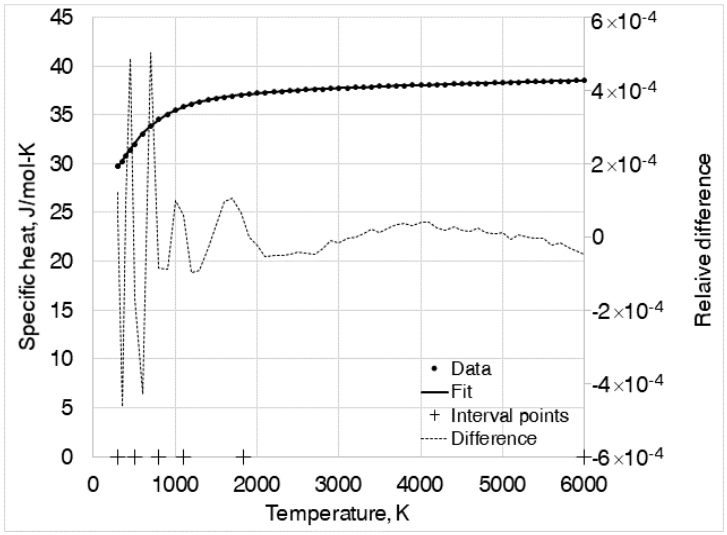


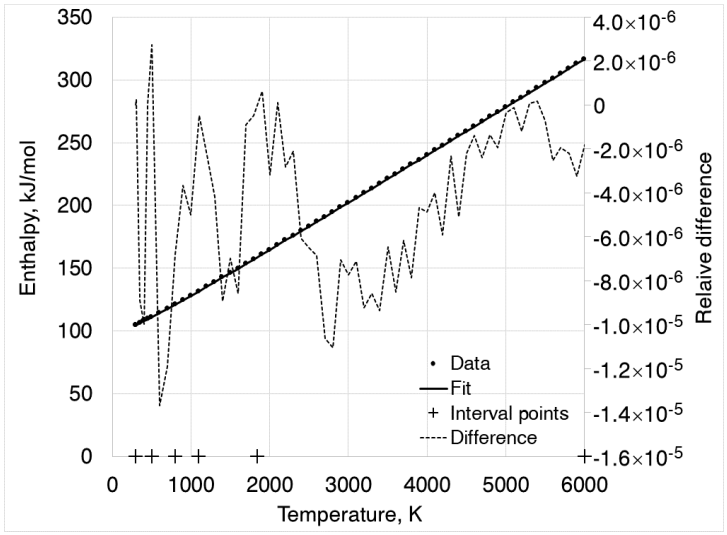


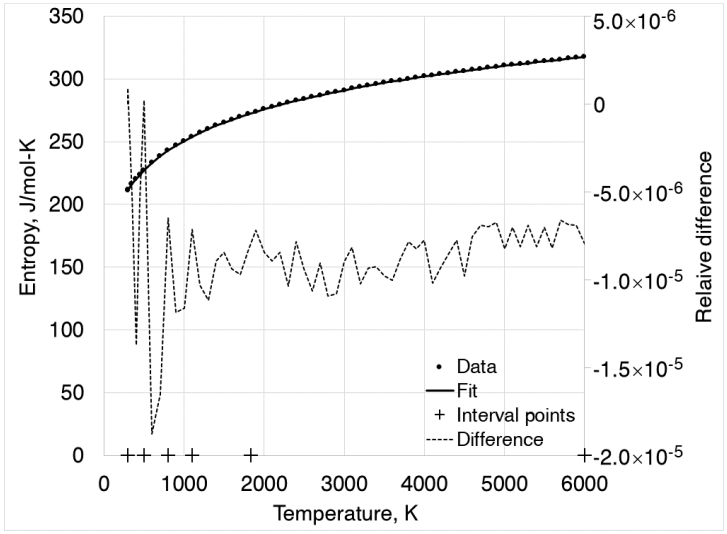


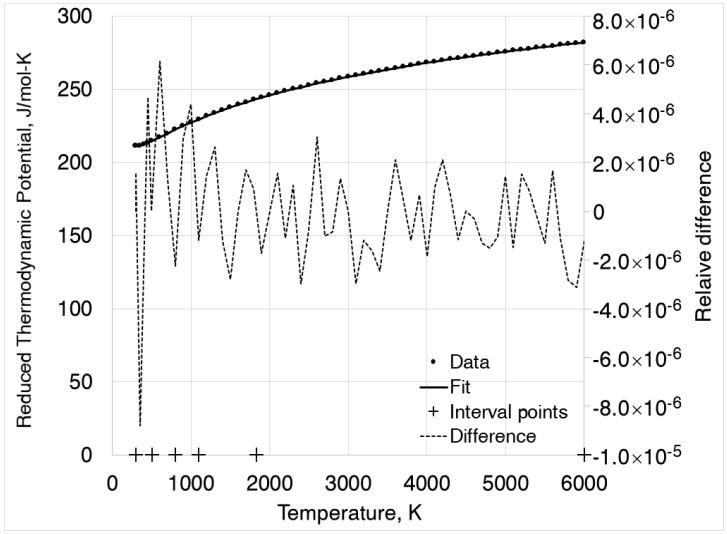
Phosphorus Nitride (PN) – N1P1(g)

MW = 44.9805 g/mol – $\Delta_f H^o$(298.15 K) = 104.78 kJ/mol – S^o(298.15 K) = 211.137 J/mol-K – p^o = 0.1 MPa



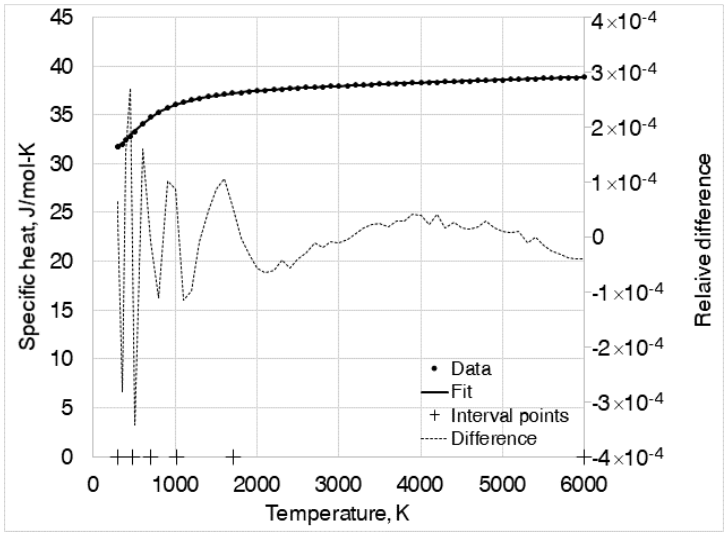


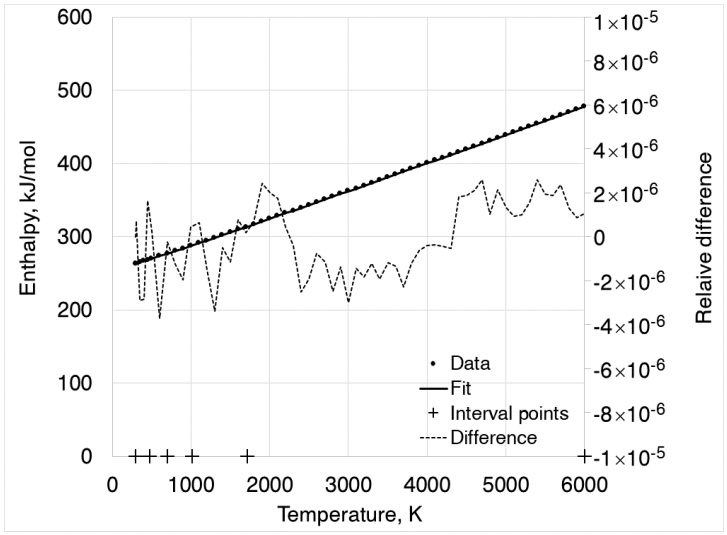


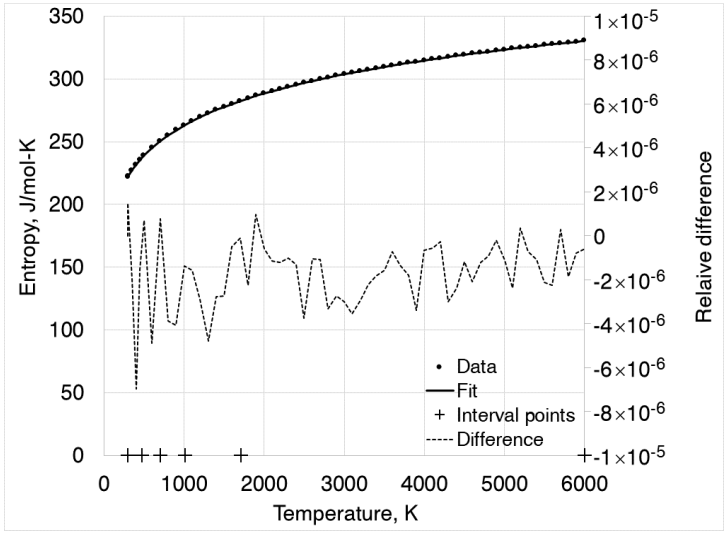


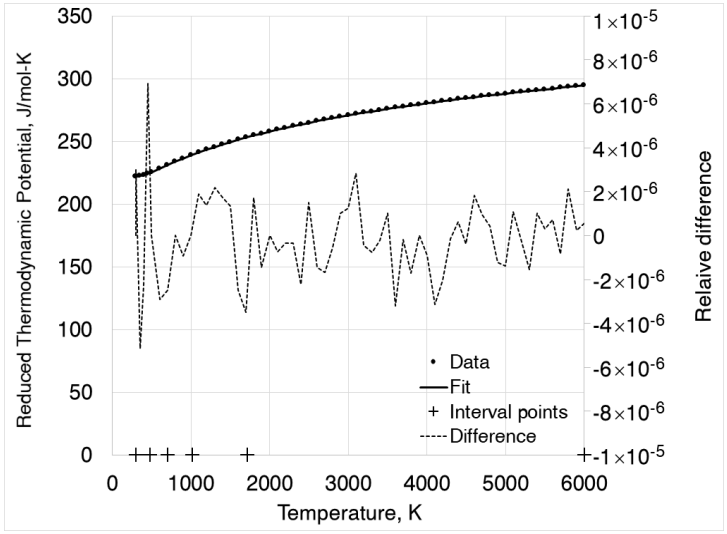
Nitrogen Sulfide (NS) – N1S1(g)

MW = 46.07 g/mol – $\Delta_f H^o$(298.15 K) = 263.592 kJ/mol – S^o(298.15 K) = 222.087 J/mol-K – p^o = 0.1 MPa



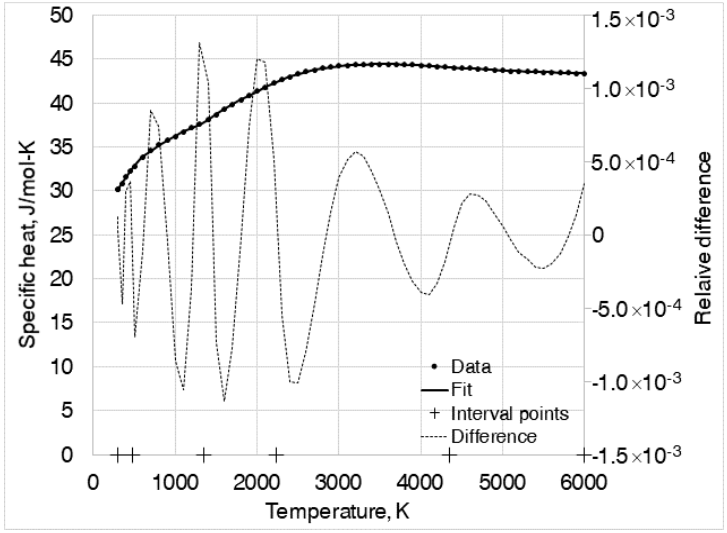


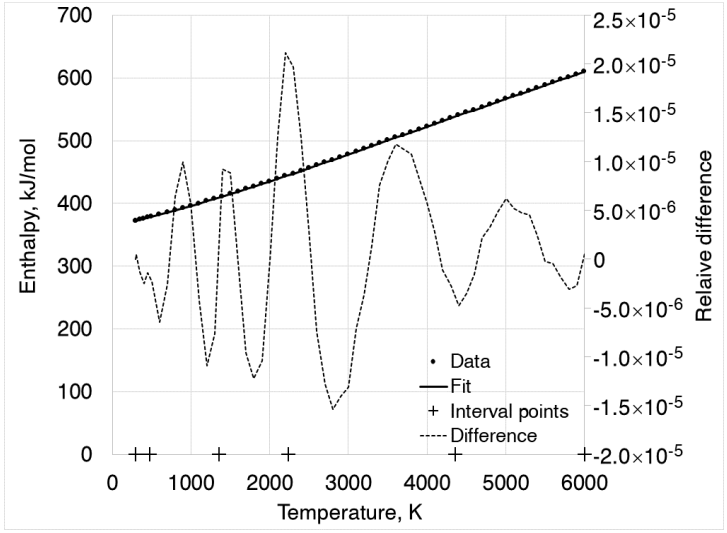


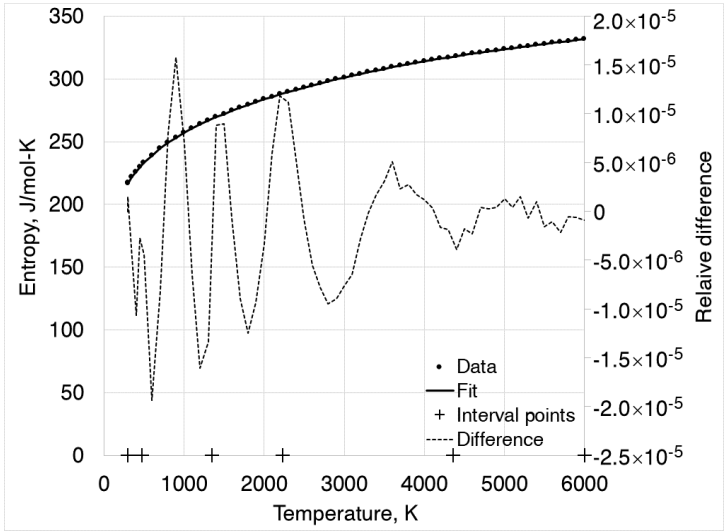


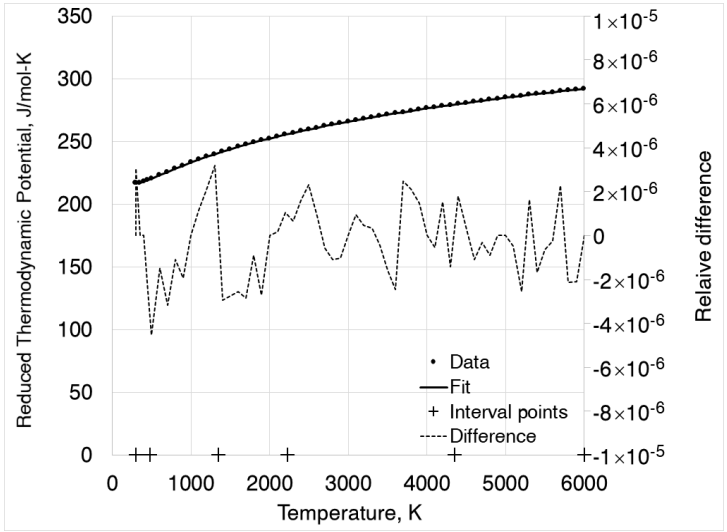
Silicon Nitride (SiN) – N1Si1(g)

MW = 42.0922 g/mol – $\Delta_f H^o$(298.15 K) = 372.376 kJ/mol – S^o(298.15 K) = 216.812 J/mol-K – p^o = 0.1 MPa



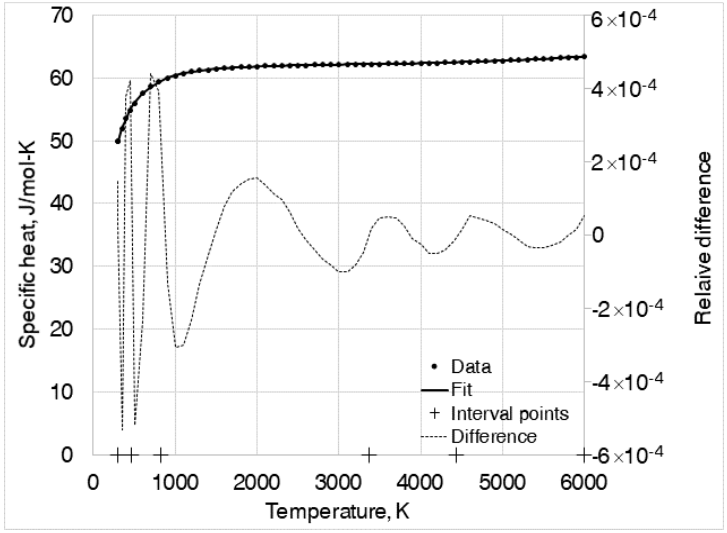


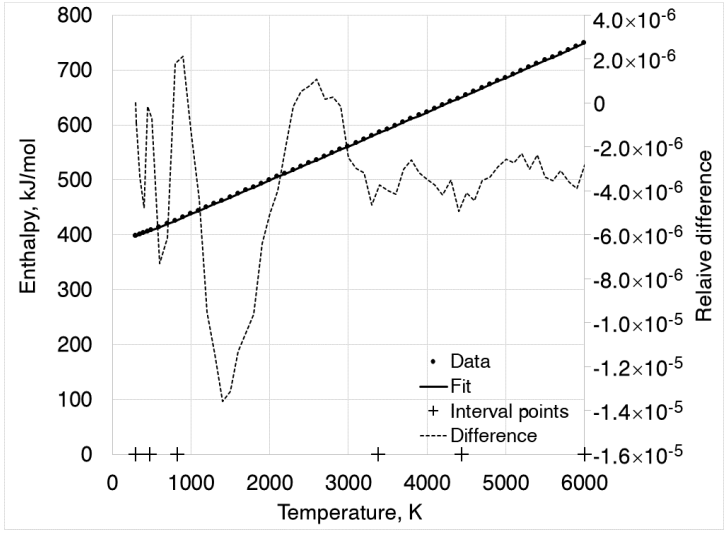


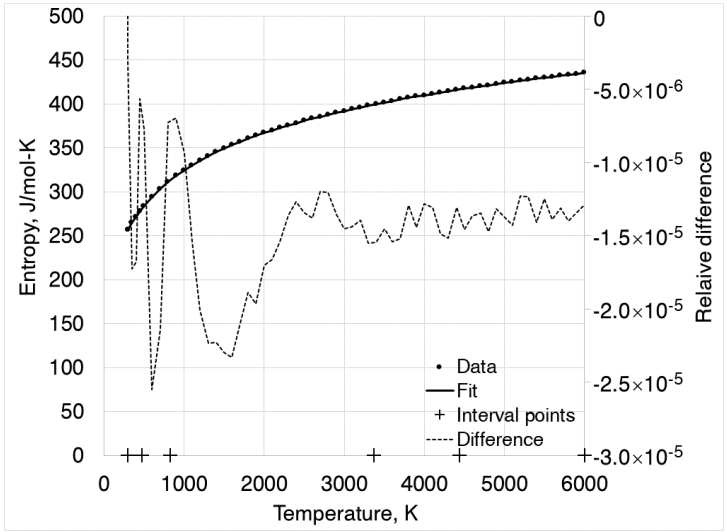


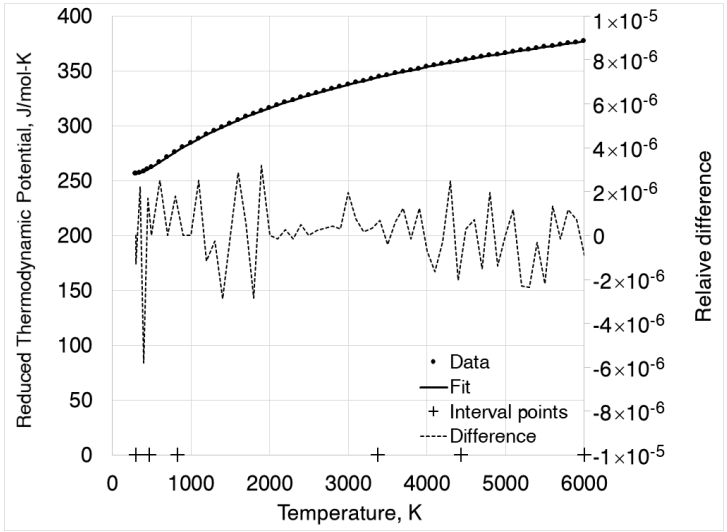
Silicon Nitride (Si_2N) – N1Si2(g)

MW = 70.1777 g/mol – $\Delta_f H^o$(298.15 K) = 397.48 kJ/mol – S^o(298.15 K) = 256.487 J/mol-K – p^o = 0.1 MPa



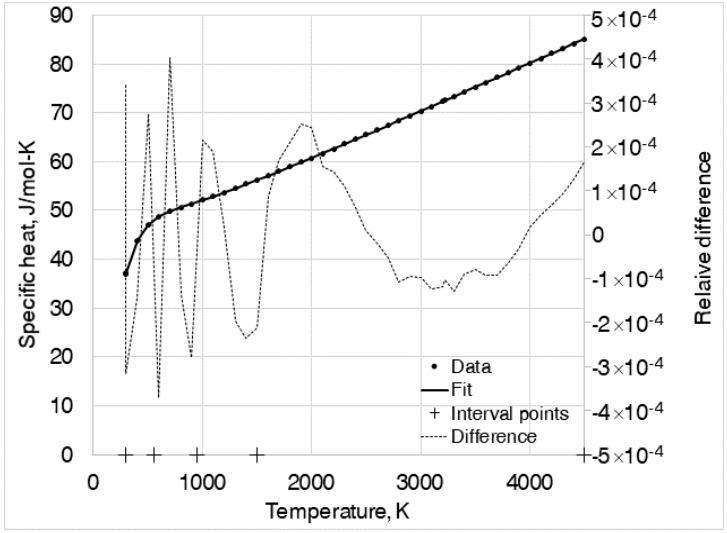


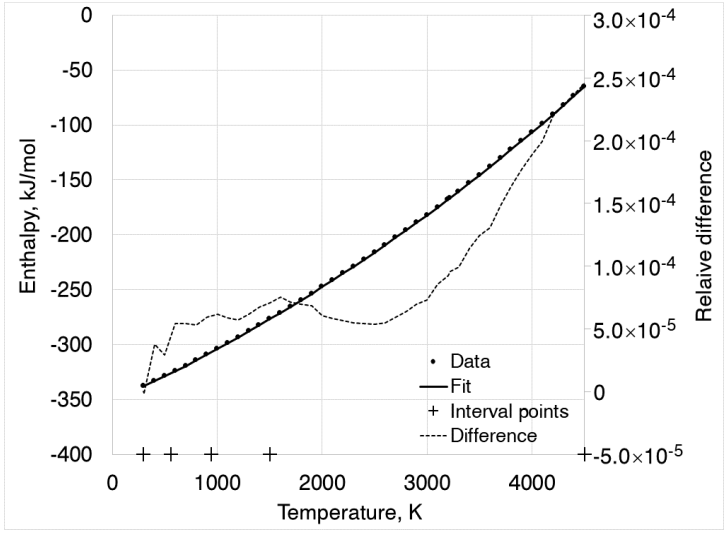


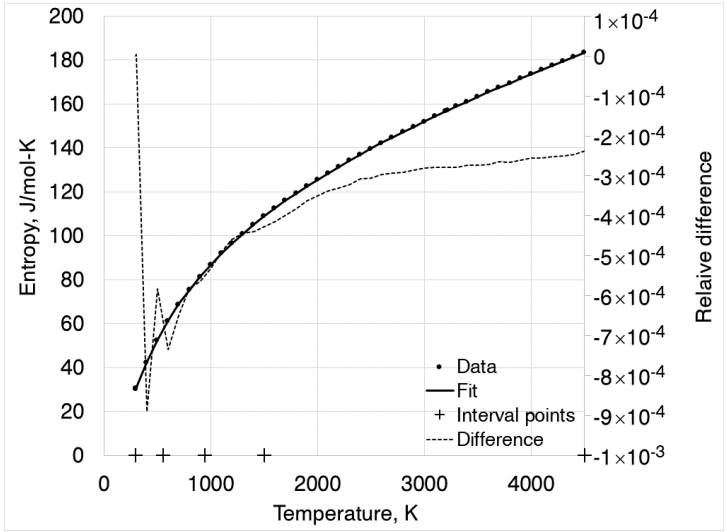


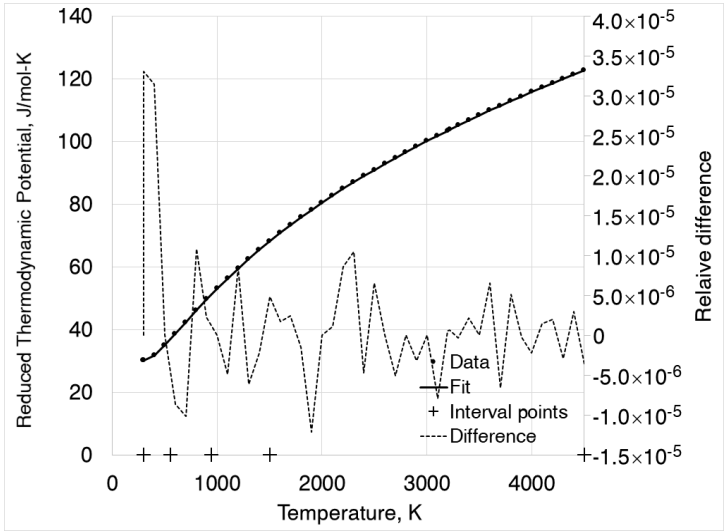
Titanium Nitride (TiN) – N1Ti1(cr)

MW = 61.89 g/mol – $\Delta_f H^o$(298.15 K) = -337.649 kJ/mol – S^o(298.15 K) = 30.235 J/mol-K – p^o = 0.1 MPa



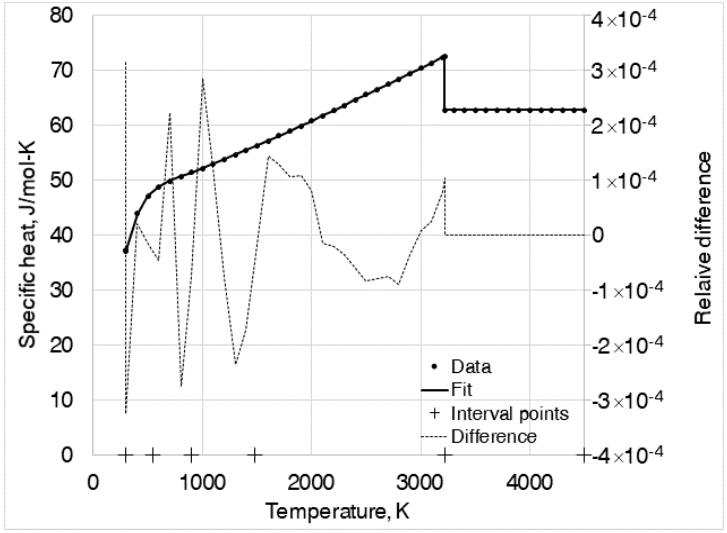


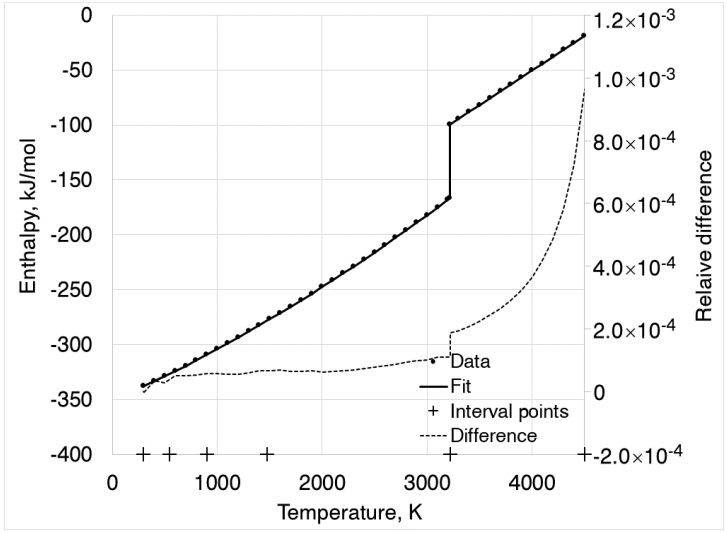


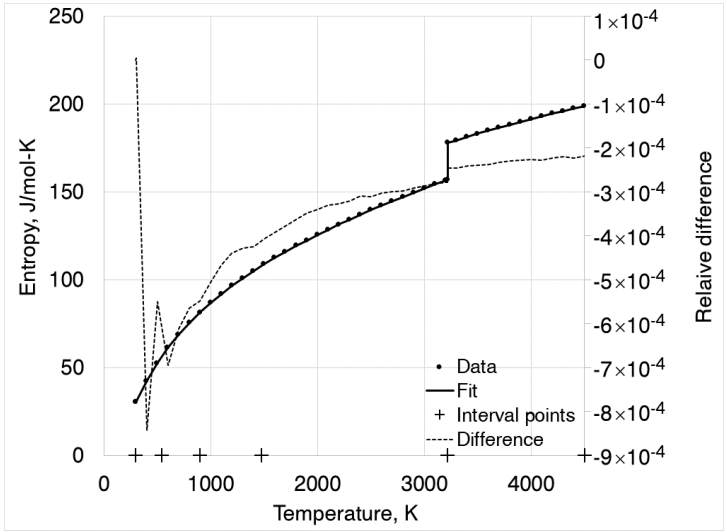


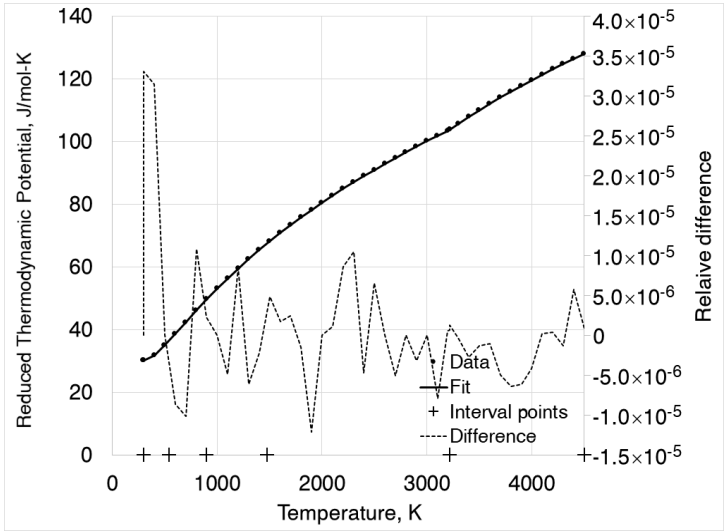
Titanium Nitride (TiN) – N1Ti1(cr,l)

MW = 61.89 g/mol – $\Delta_f H^o$(298.15 K) = -337.649 kJ/mol – S^o(298.15 K) = 30.235 J/mol-K – p^o = 0.1 MPa



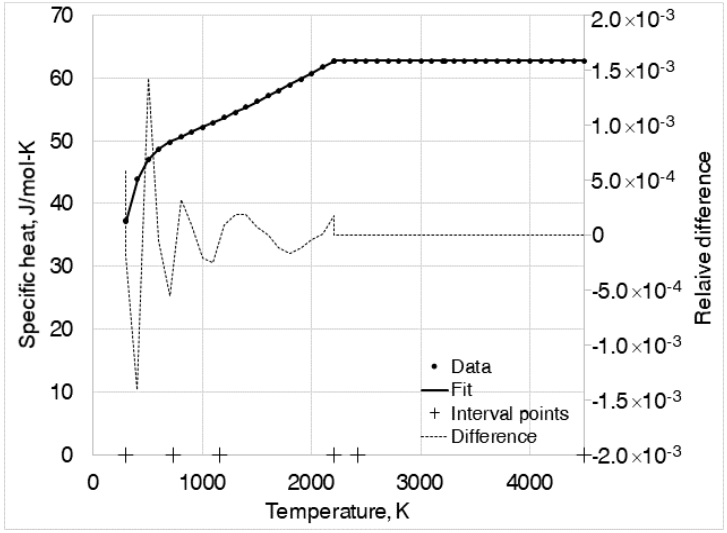


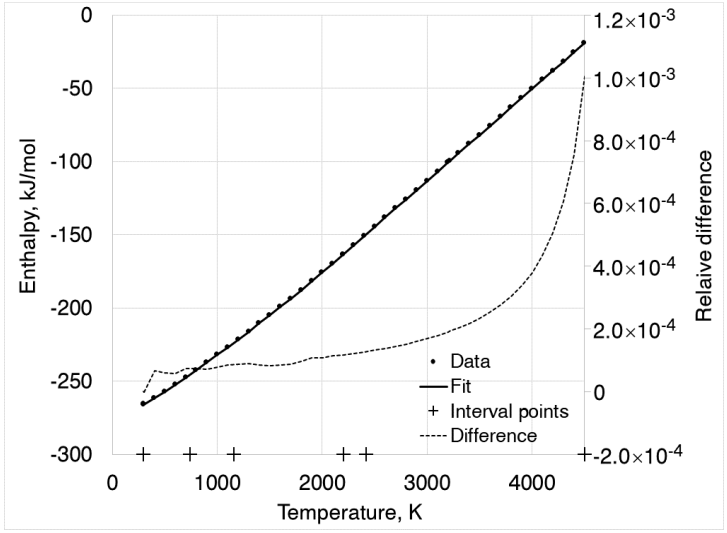


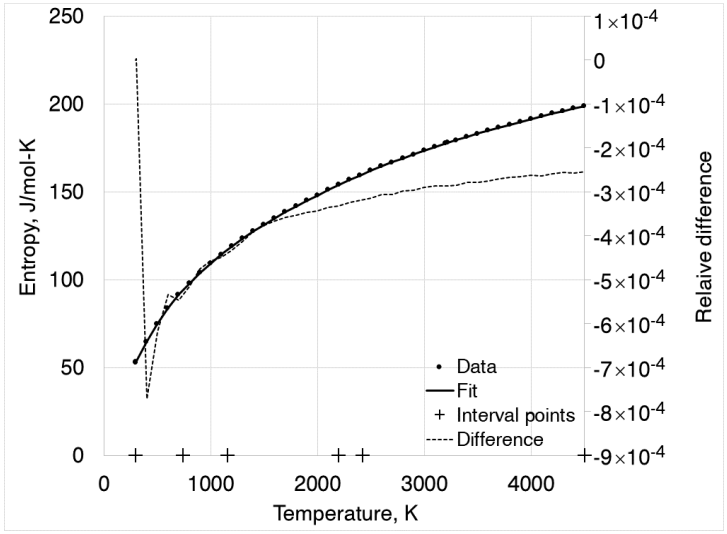


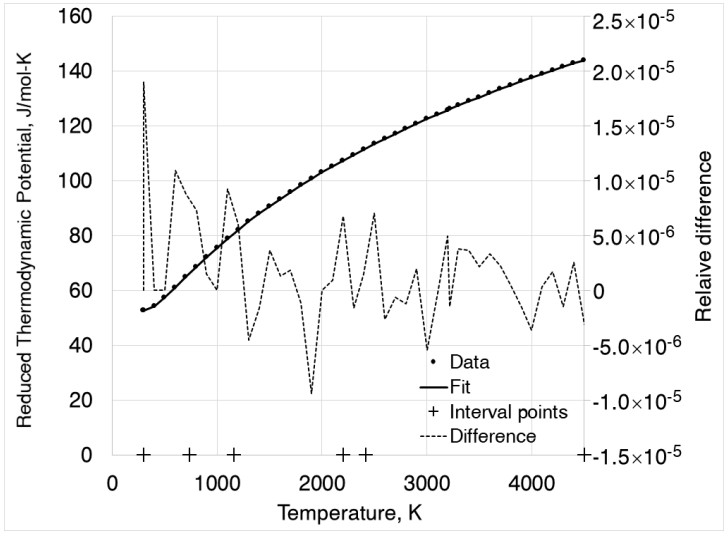
Titanium Nitride (TiN) – N1Ti1(l)

MW = 61.89 g/mol – $\Delta_f H^o$(298.15 K) = -265.813 kJ/mol – S^o(298.15 K) = 52.733 J/mol-K – p^o = 0.1 MPa



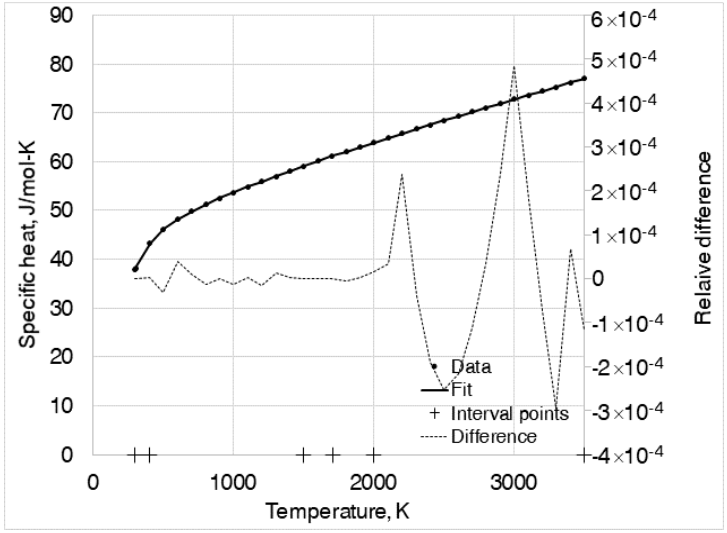


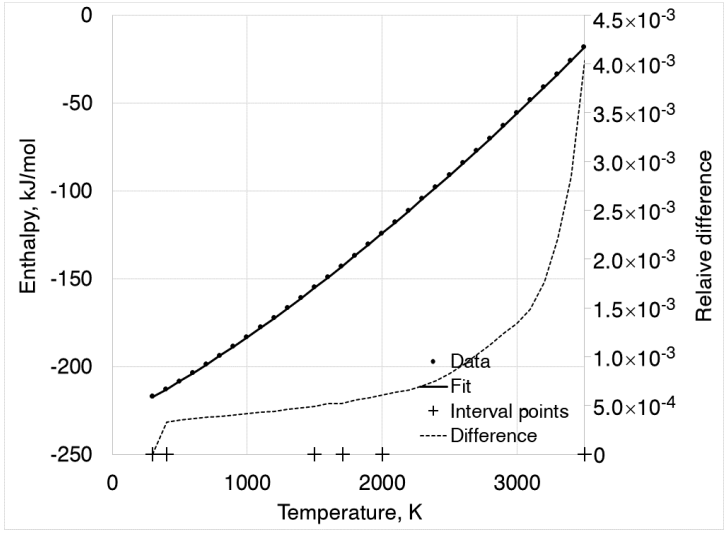


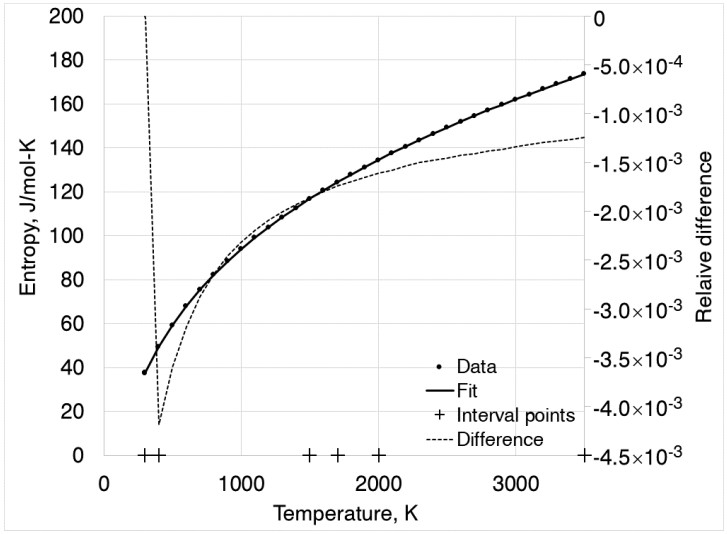


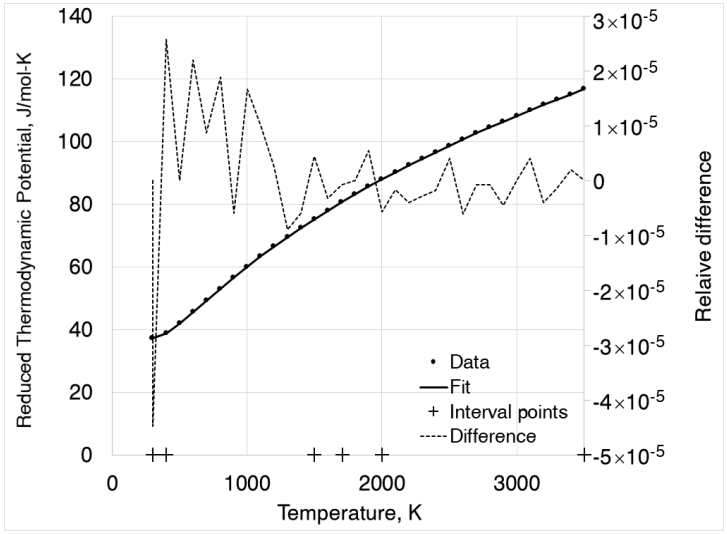
Vanadium Nitride (VN) – N1V1(cr)

MW = 64.9482 g/mol – $\Delta_f H^o$(298.15 K) = -217.15 kJ/mol – S^o(298.15 K) = 37.264 J/mol-K – p^o = 0.1 MPa



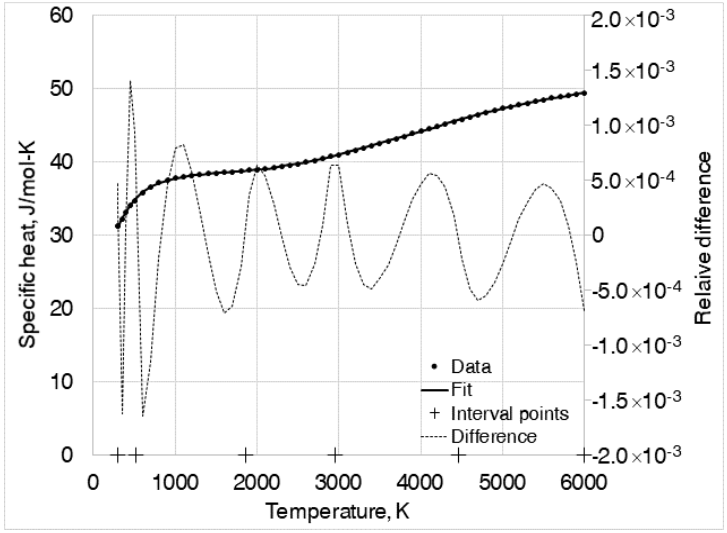


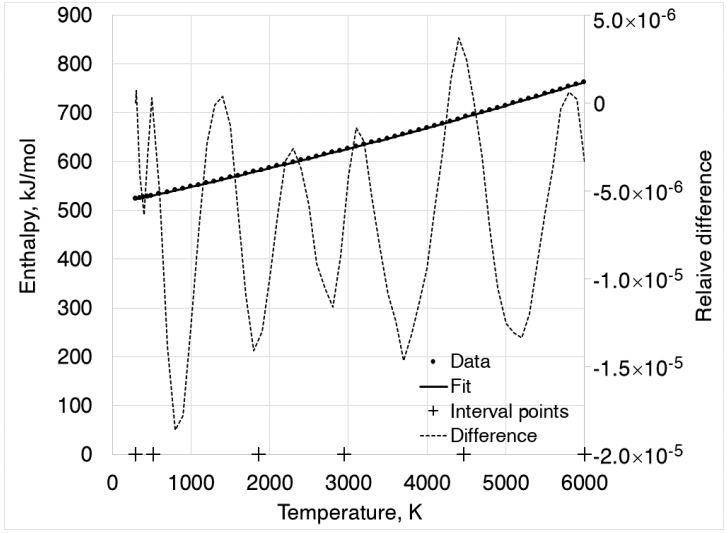


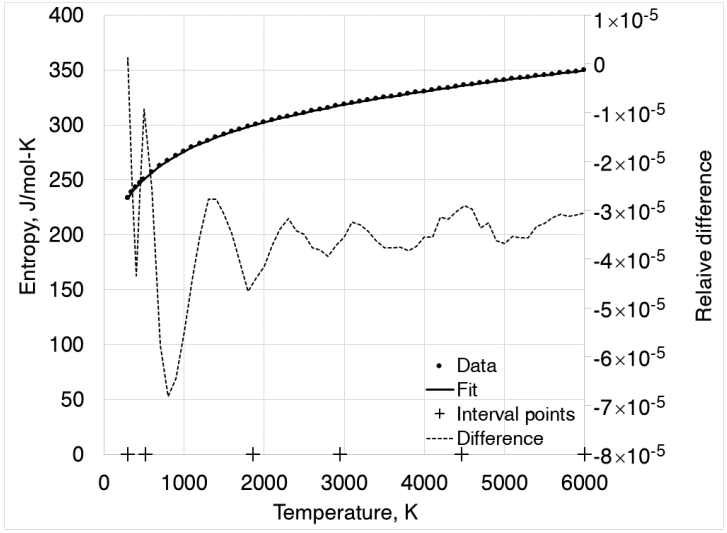


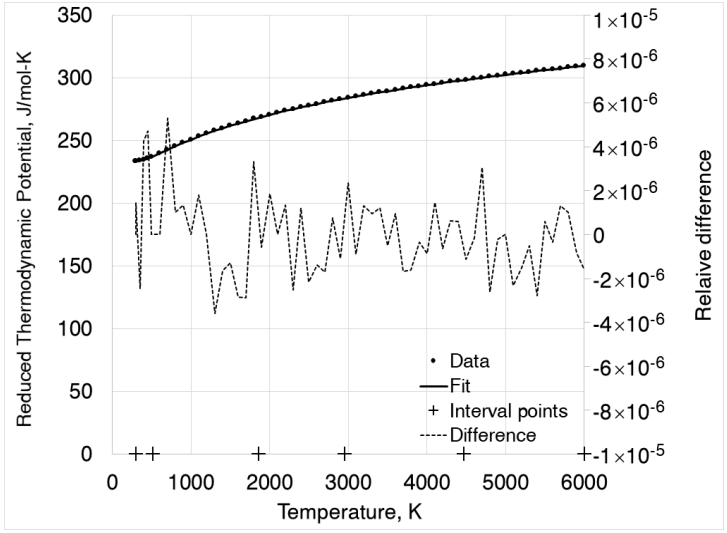
Vanadium Nitride (VN) – N1V1(g)

MW = 64.9482 g/mol – $\Delta_f H^o$(298.15 K) = 523 kJ/mol – S^o(298.15 K) = 233.384 J/mol-K – p^o = 0.1 MPa



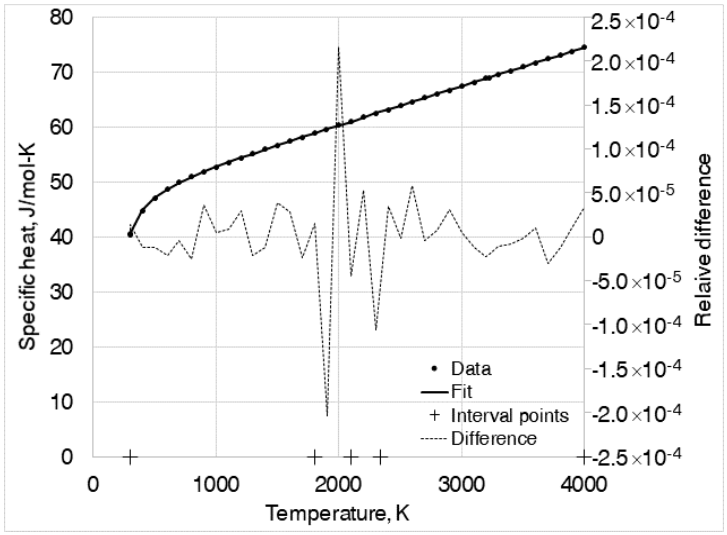


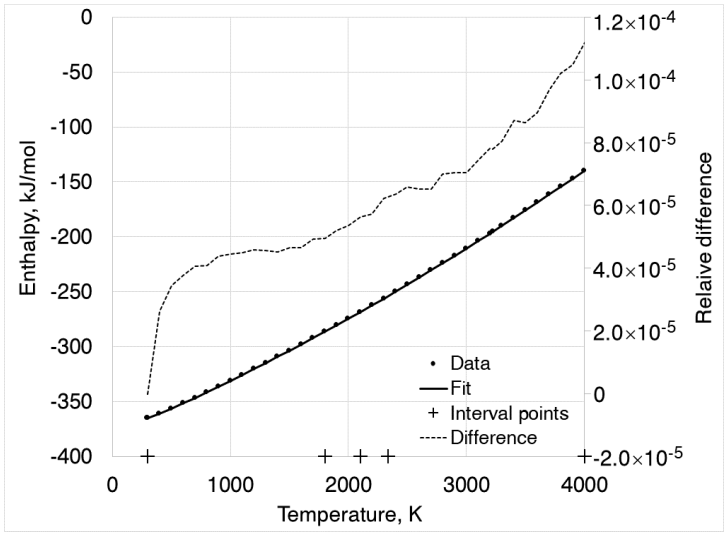


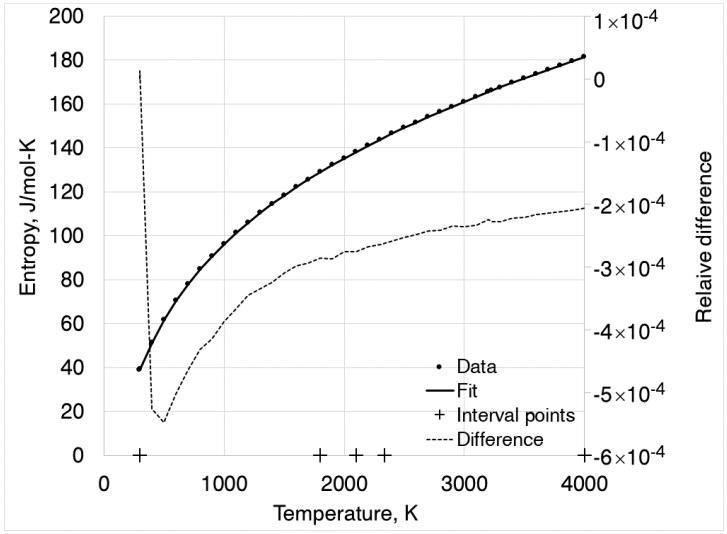


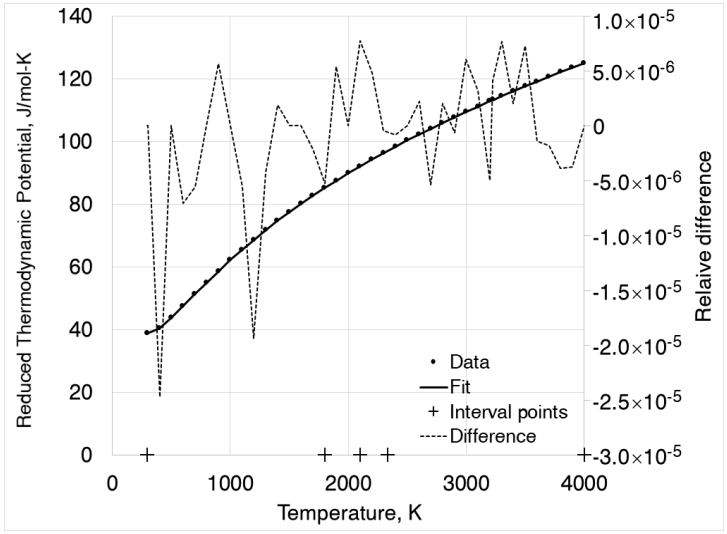
Zirconium Nitride (ZrN) – N1Zr1(cr)

MW = 105.23 g/mol – $\Delta_f H^o$(298.15 K) = -365.263 kJ/mol – S^o(298.15 K) = 38.861 J/mol-K – p^o = 0.1 MPa



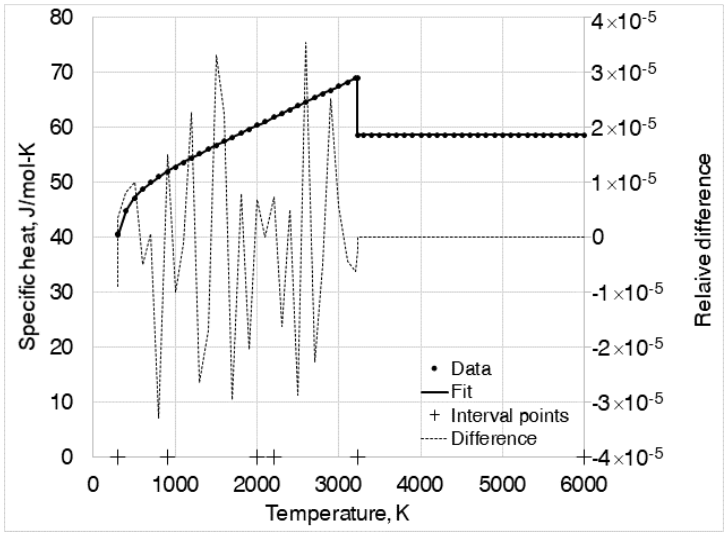


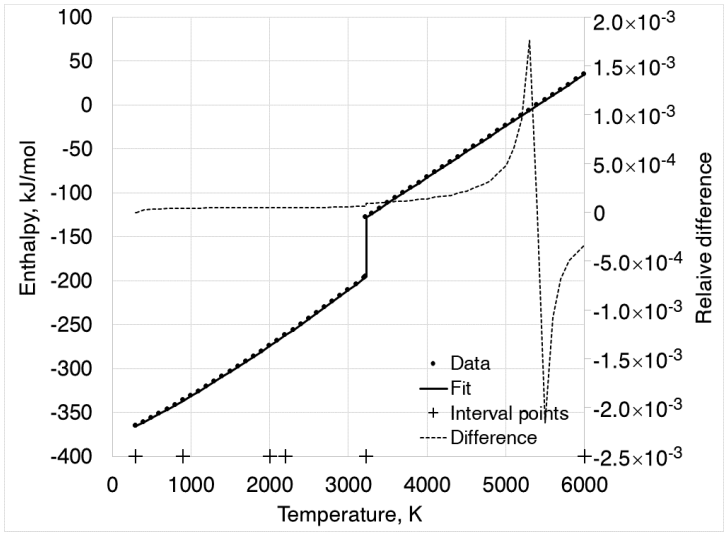


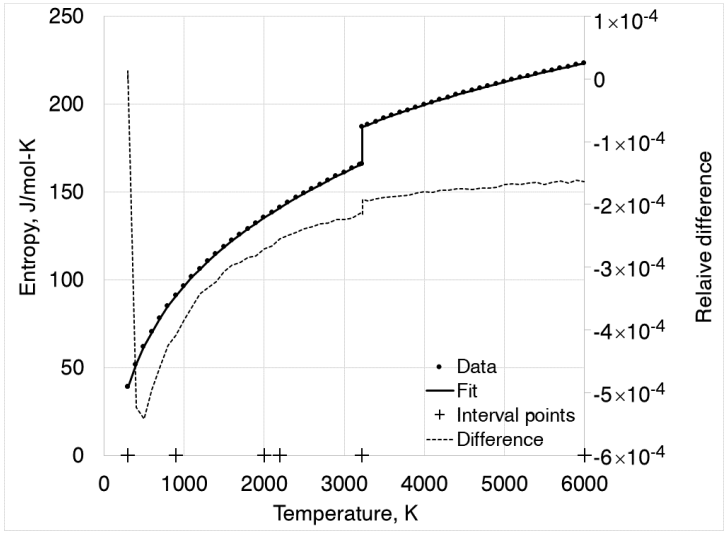


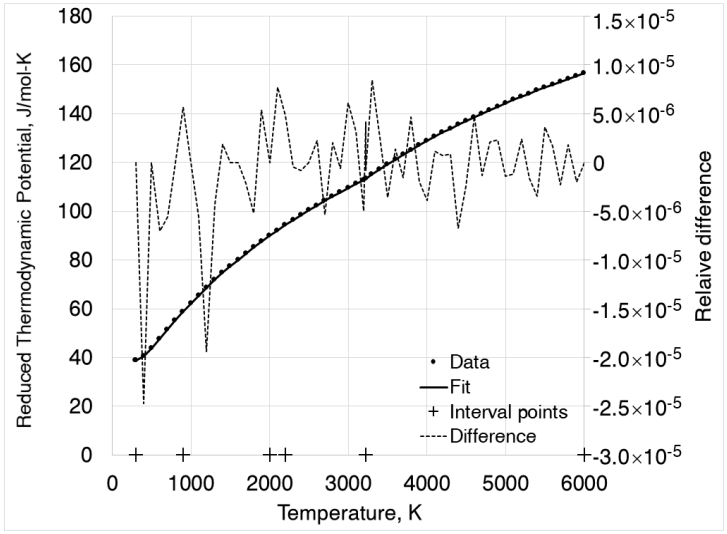
Zirconium Nitride (ZrN) – N1Zr1(cr,l)

MW = 105.23 g/mol – $\Delta_f H^o$(298.15 K) = -365.263 kJ/mol – S^o(298.15 K) = 38.861 J/mol-K – p^o = 0.1 MPa



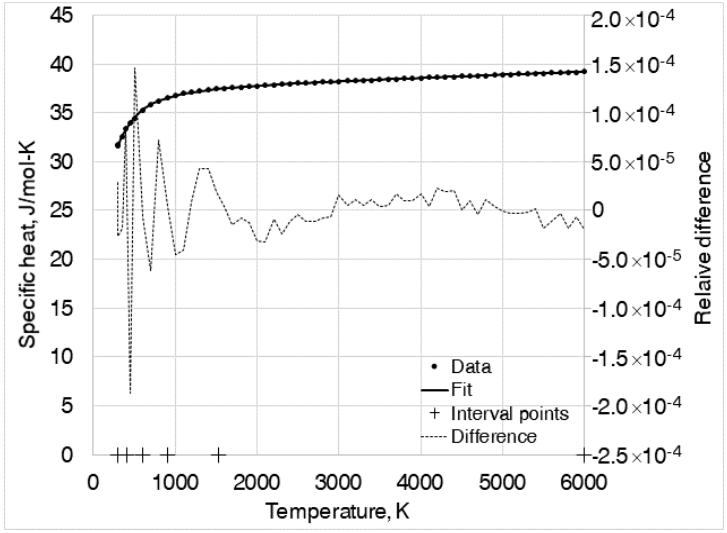


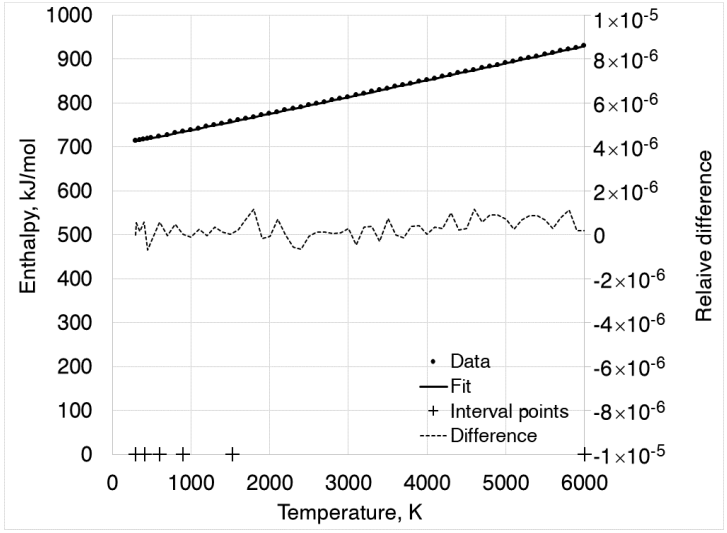


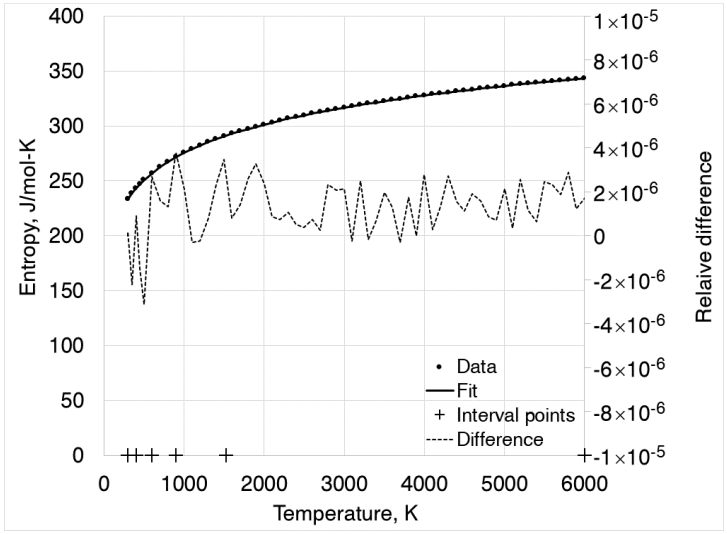


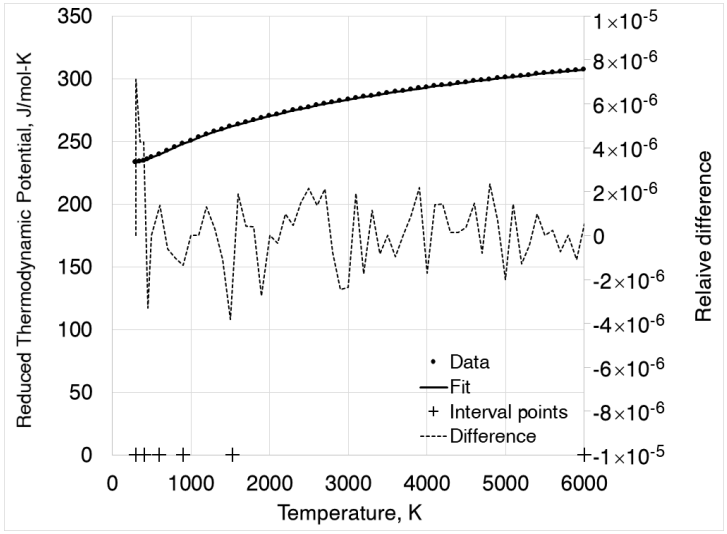
Zirconium Nitride (ZrN) – N1Zr1(g)

MW = 105.23 g/mol – $\Delta_f H^o$(298.15 K) = 713.372 kJ/mol – S^o(298.15 K) = 233.487 J/mol-K – p^o = 0.1 MPa



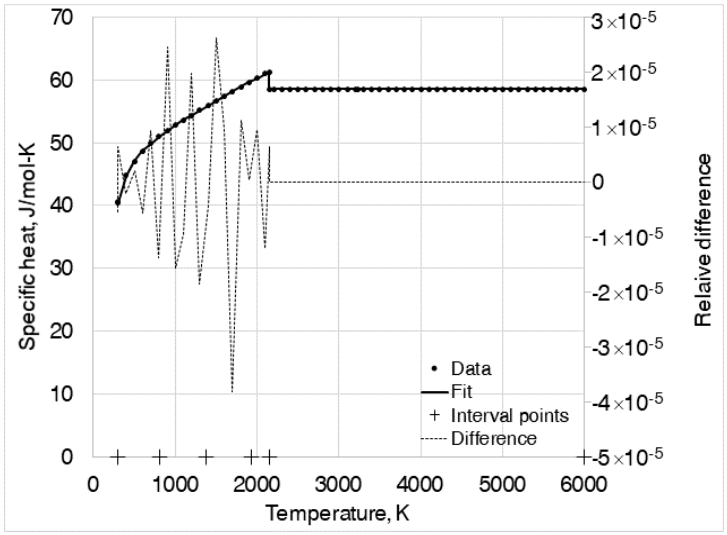


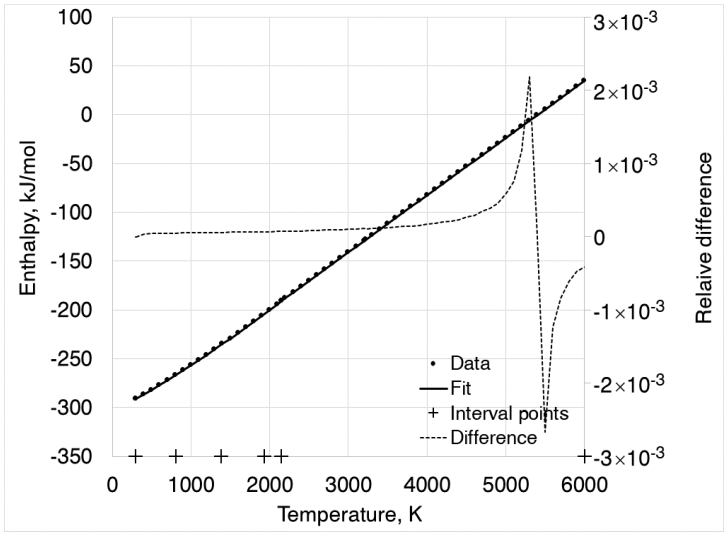


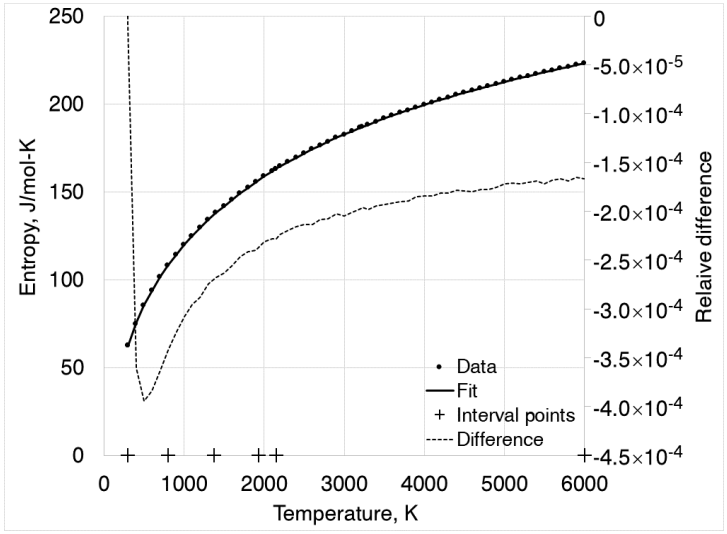


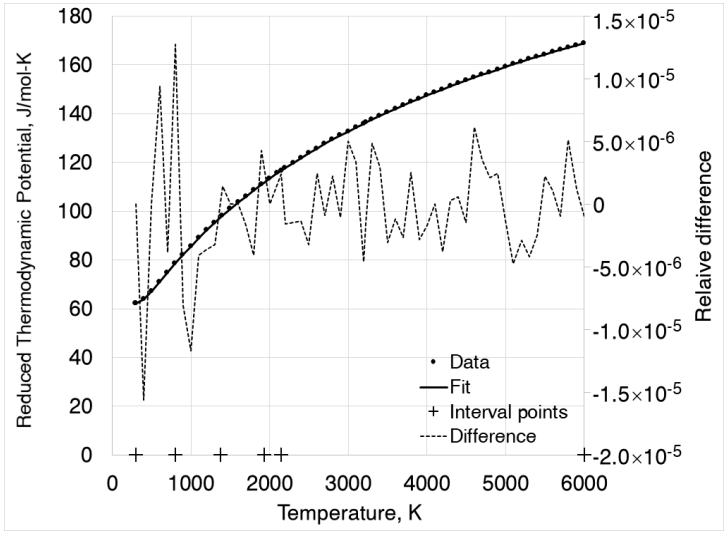
Zirconium Nitride (ZrN) – N1Zr1(l)

MW = 105.23 g/mol – $\Delta_f H^o$(298.15 K) = -290.749 kJ/mol – S^o(298.15 K) = 62.342 J/mol-K – p^o = 0.1 MPa



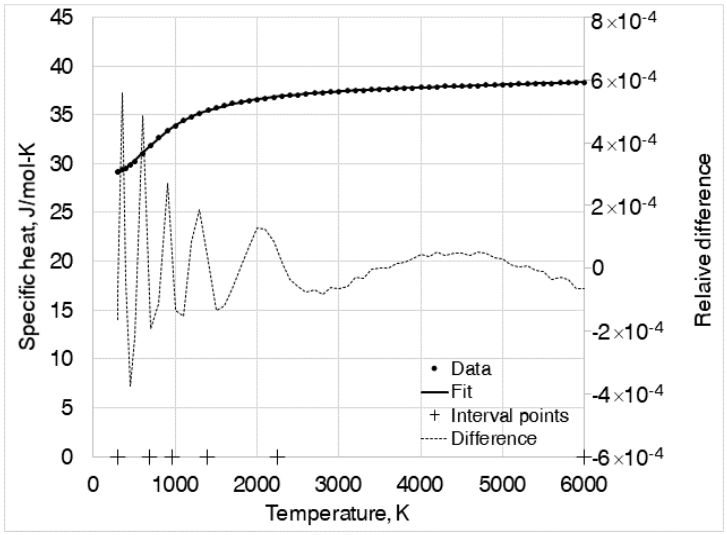


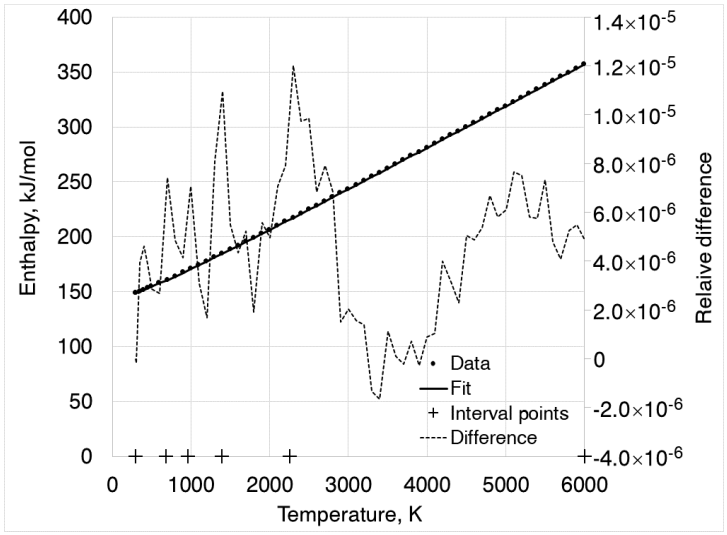


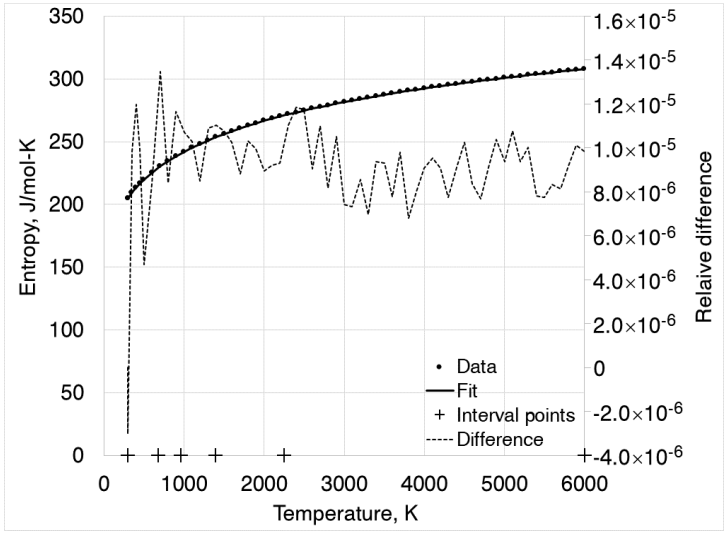


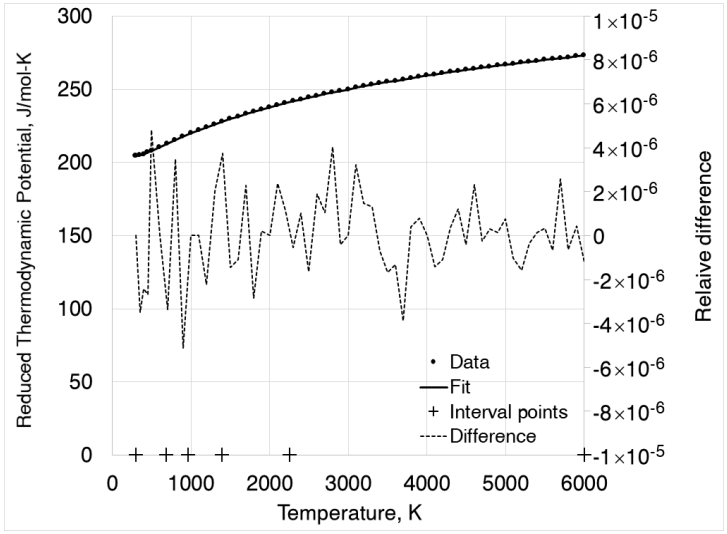
Nitrogen, Ion (N_2^-) – N2-(g)

MW = 28.0139 g/mol – $\Delta_f H^o$(298.15 K) = 148.615 kJ/mol – S^o(298.15 K) = 204.536 J/mol-K – p^o = 0.1 MPa



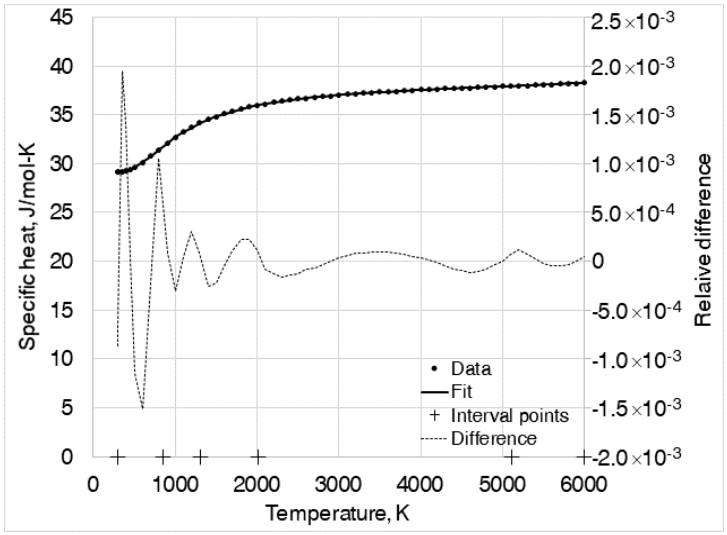


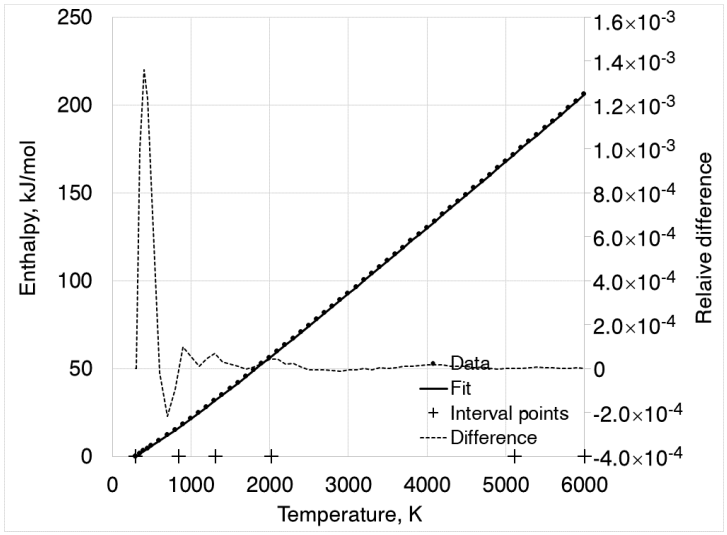


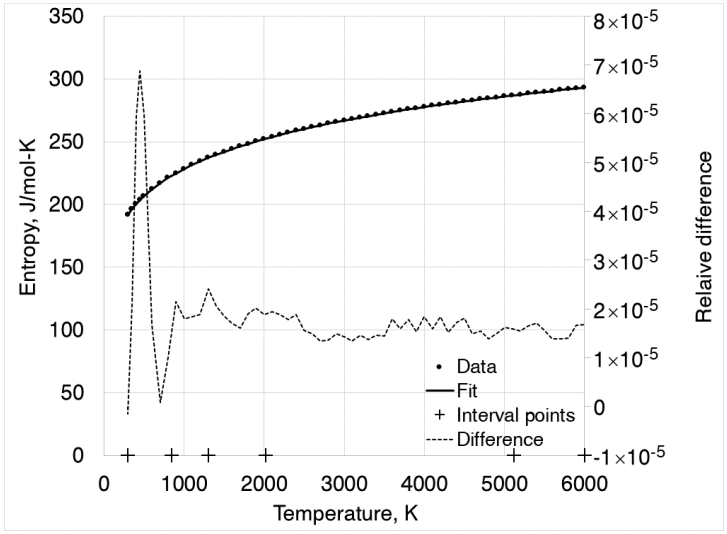


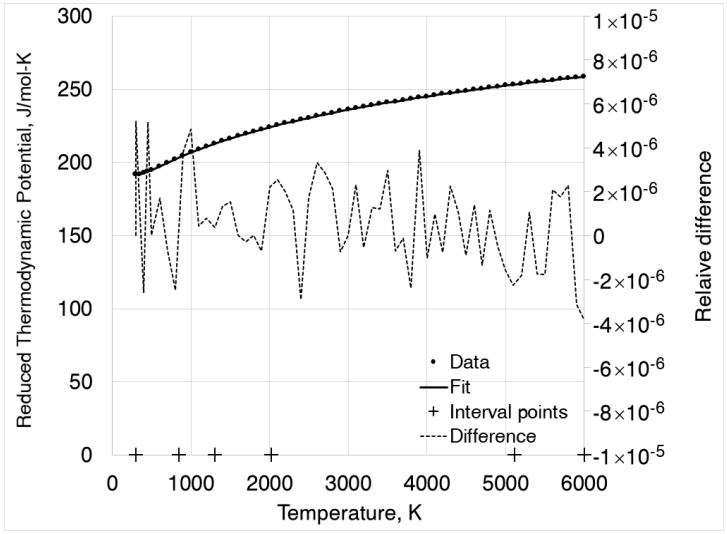
Nitrogen (N_2) – N2(ref)

MW = 28.0134 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 191.609 J/mol-K – p^o = 0.1 MPa



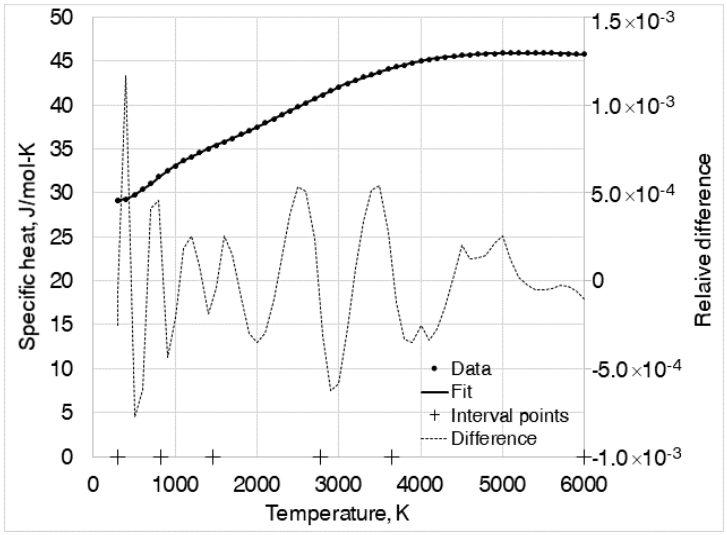


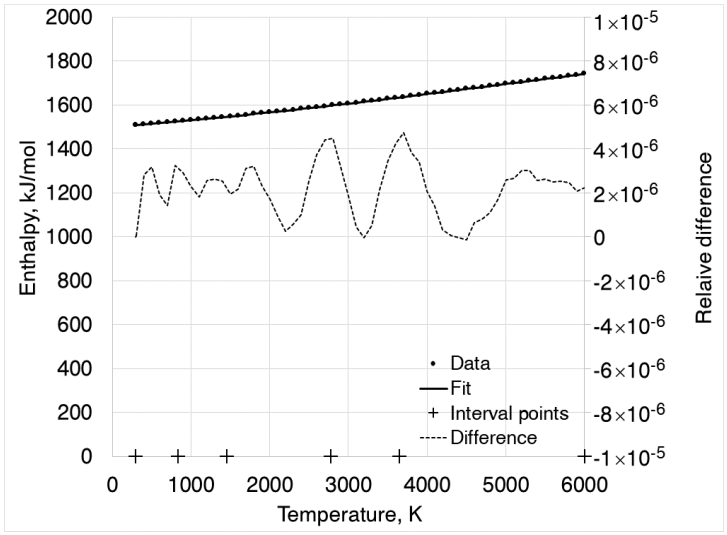


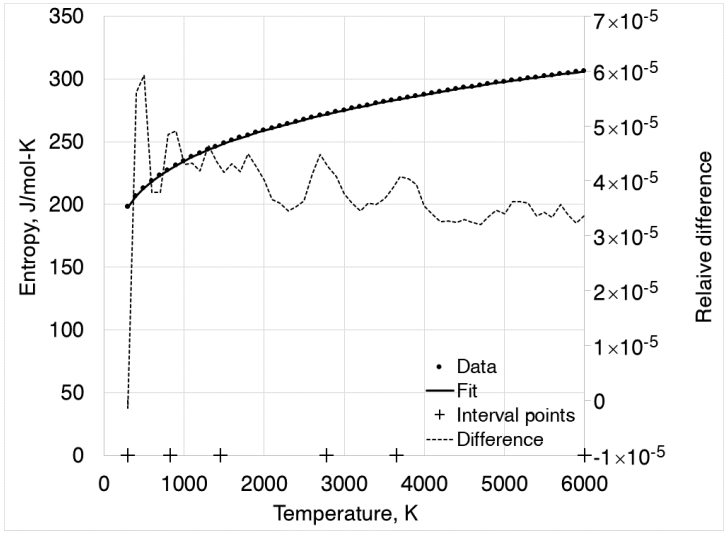


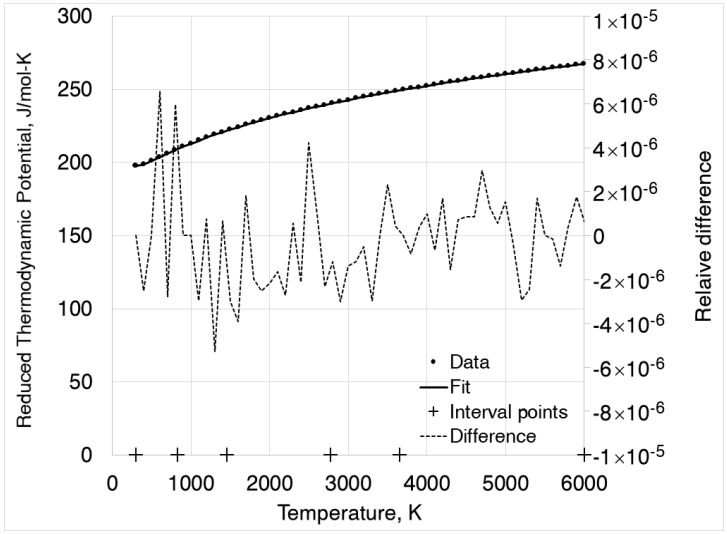
Nitrogen, Ion (N_2^+) – N2+(g)

MW = 28.0129 g/mol – $\Delta_f H^o$(298.15 K) = 1509.499 kJ/mol – S^o(298.15 K) = 197.66 J/mol-K – p^o = 0.1 MPa



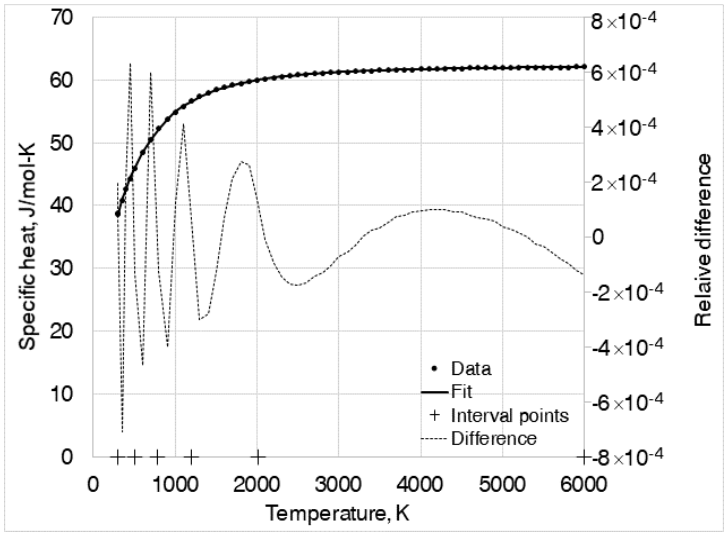


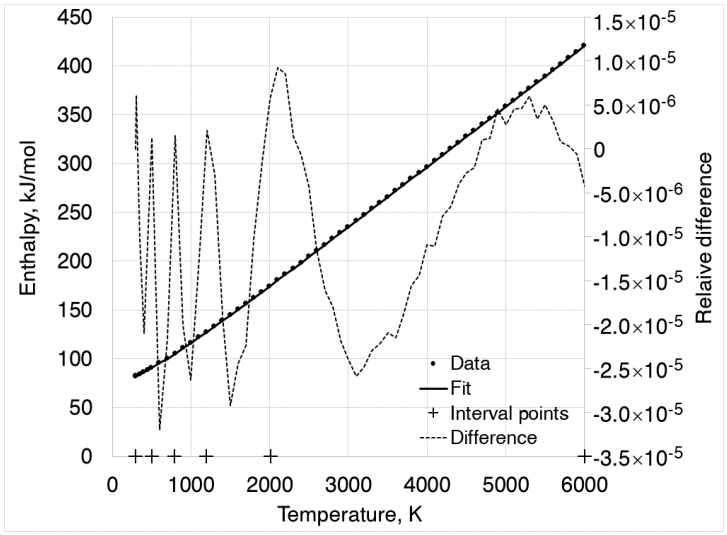


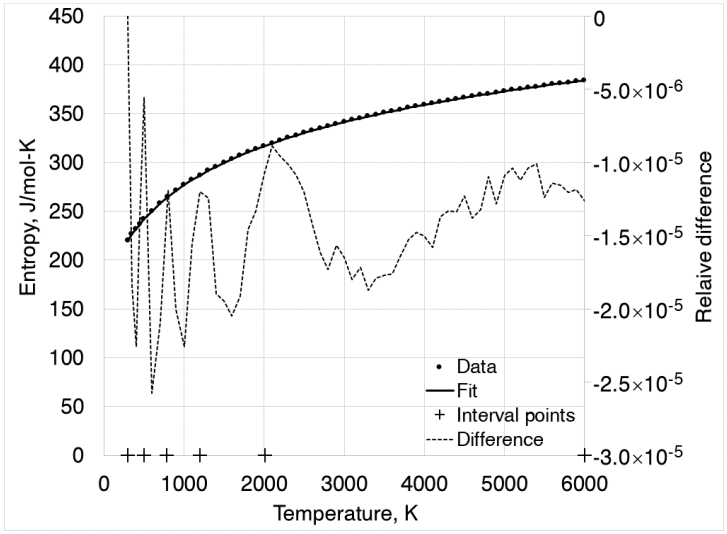


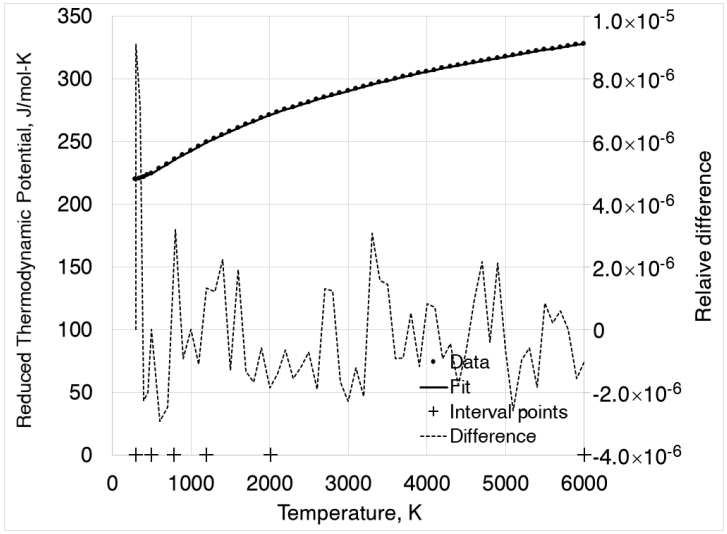
Nitrogen Oxide (N_2O) – N2O1(g)

MW = 44.0128 g/mol – $\Delta_f H^o$(298.15 K) = 82.048 kJ/mol – S^o(298.15 K) = 219.957 J/mol-K – p^o = 0.1 MPa



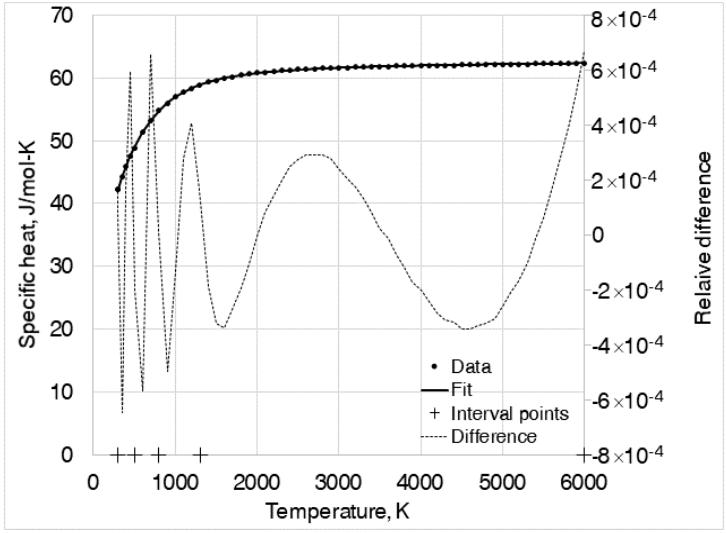


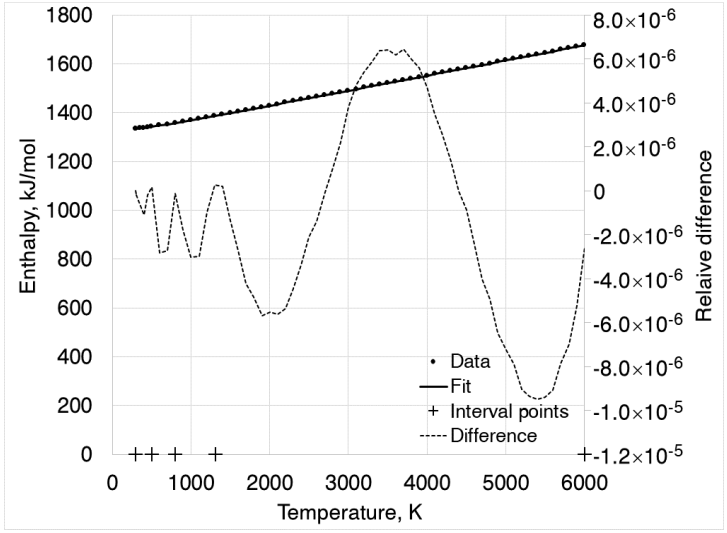


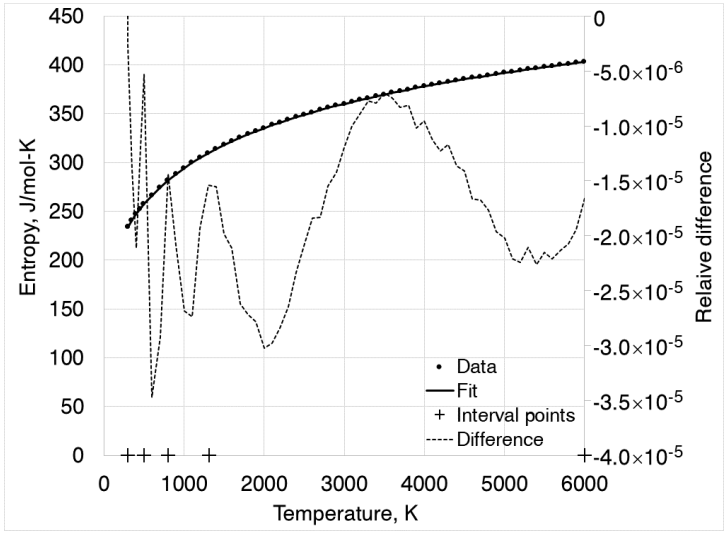


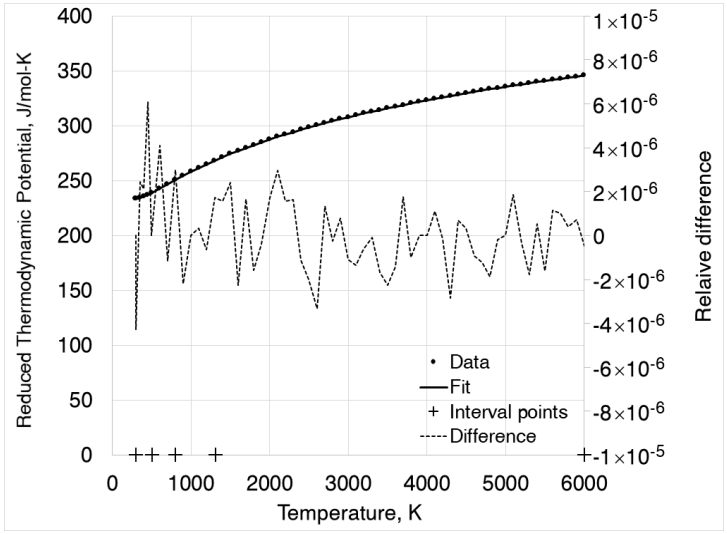
Nitrogen Oxide, Ion (N_2O^+) – N2O1+(g)

MW = 44.0123 g/mol – $\Delta_f H^o$(298.15 K) = 1333.399 kJ/mol – S^o(298.15 K) = 233.856 J/mol-K – p^o = 0.1 MPa



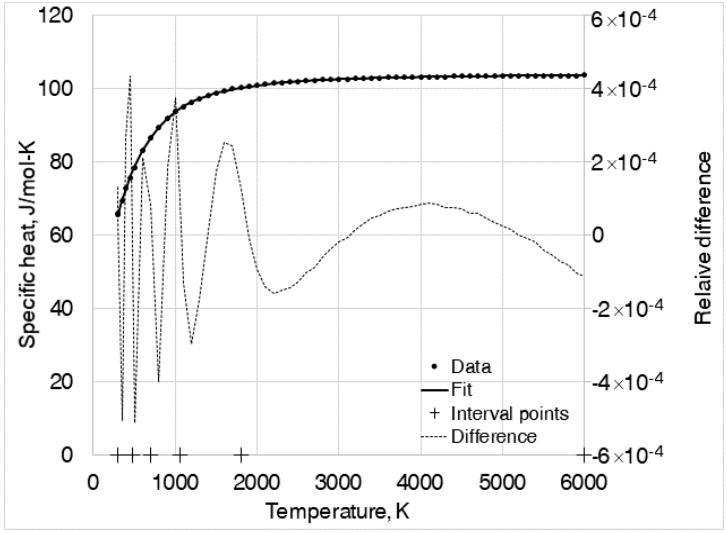


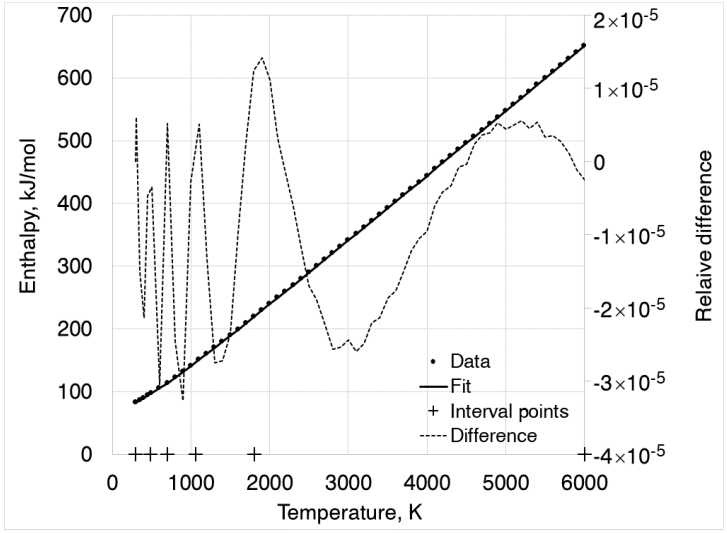


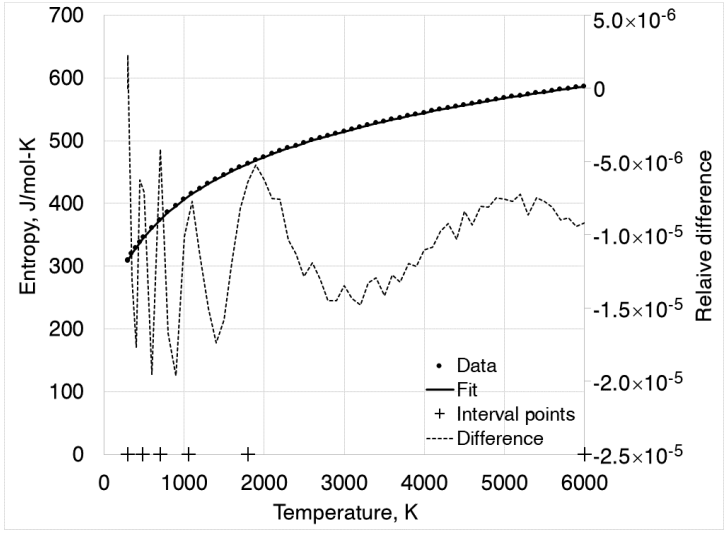


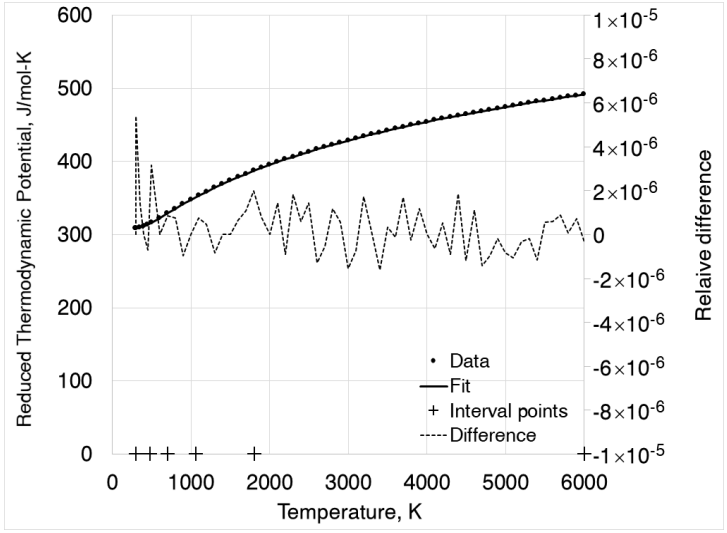
Nitrogen Oxide (N_2O_3) – N2O3(g)

MW = 76.0116 g/mol – $\Delta_f H^o$(298.15 K) = 82.843 kJ/mol – S^o(298.15 K) = 308.539 J/mol-K – p^o = 0.1 MPa



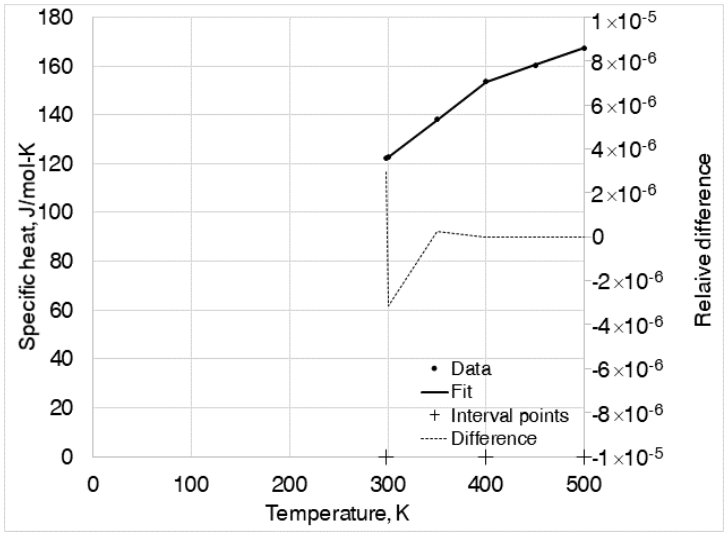


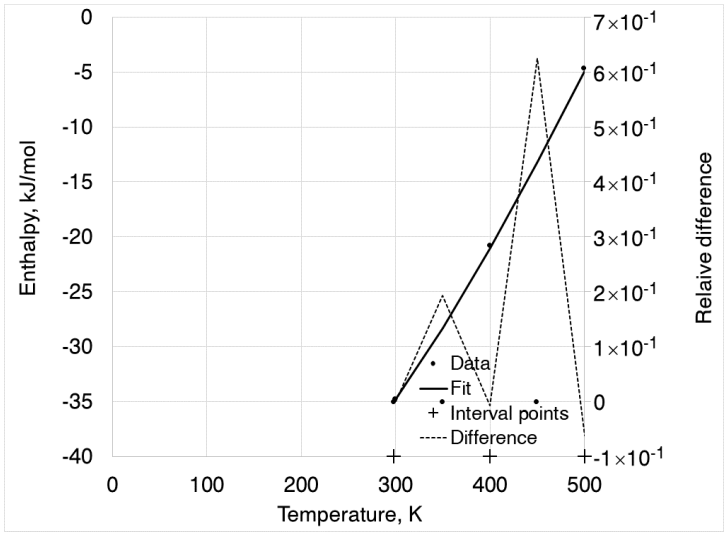


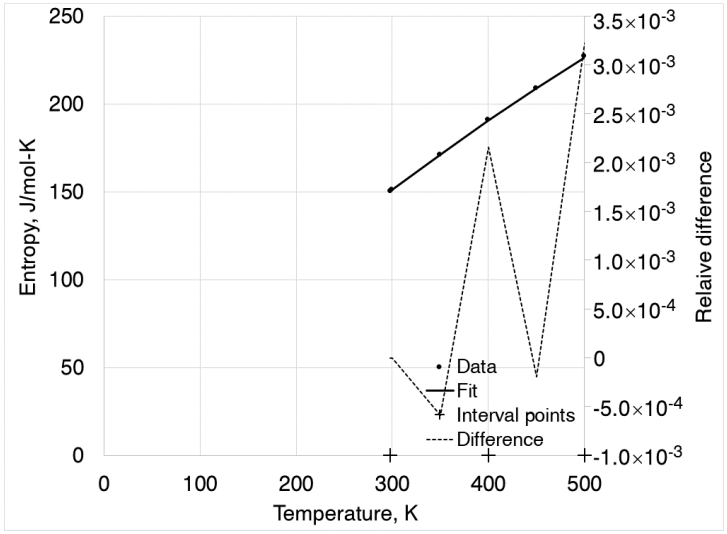


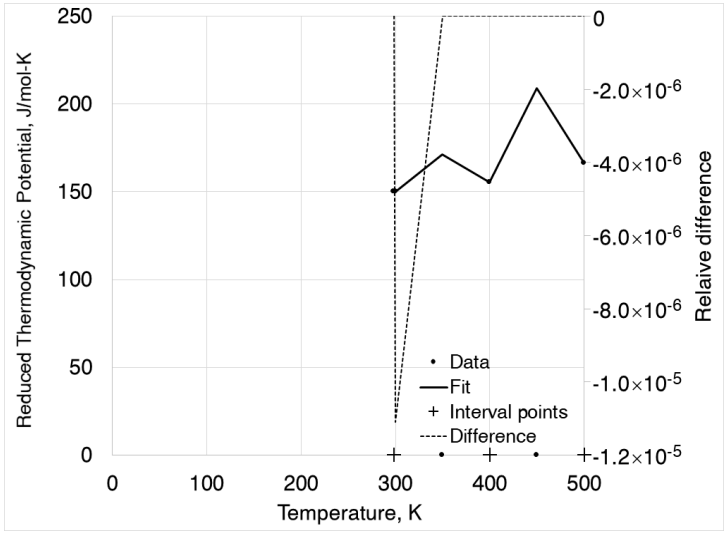
Nitrogen Oxide (N_2O_4) – N2O4(cr)

MW = 92.011 g/mol – $\Delta_f H^o$(298.15 K) = -35.046 kJ/mol – S^o(298.15 K) = 150.29 J/mol-K – p^o = 0.1 MPa



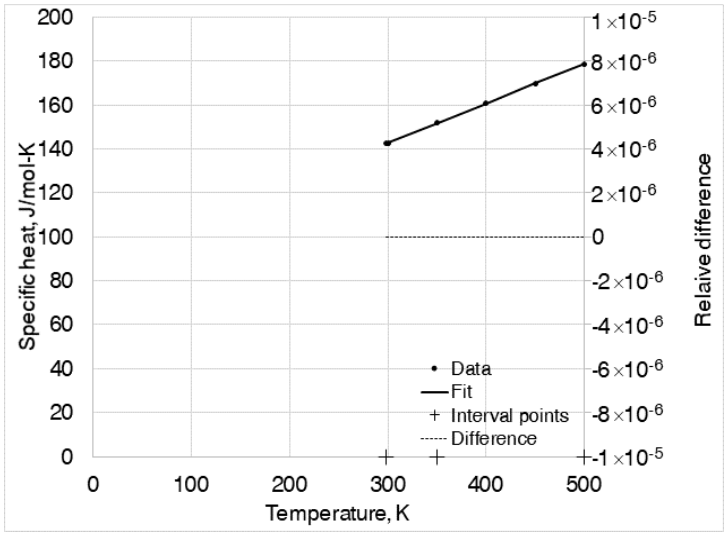


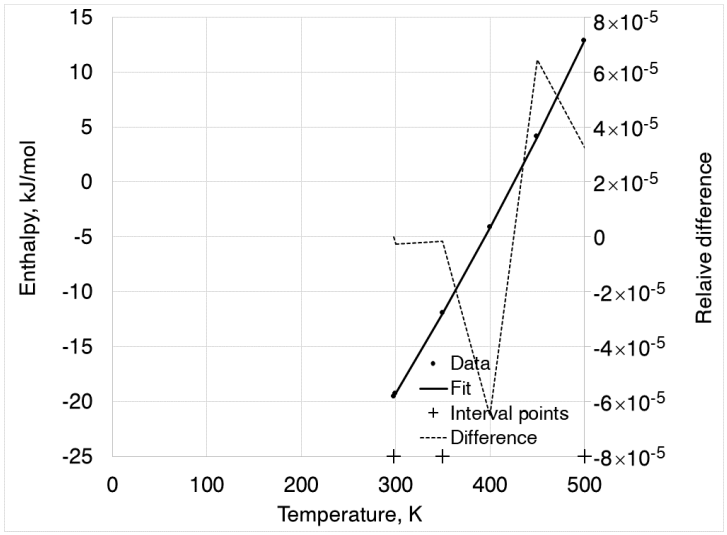


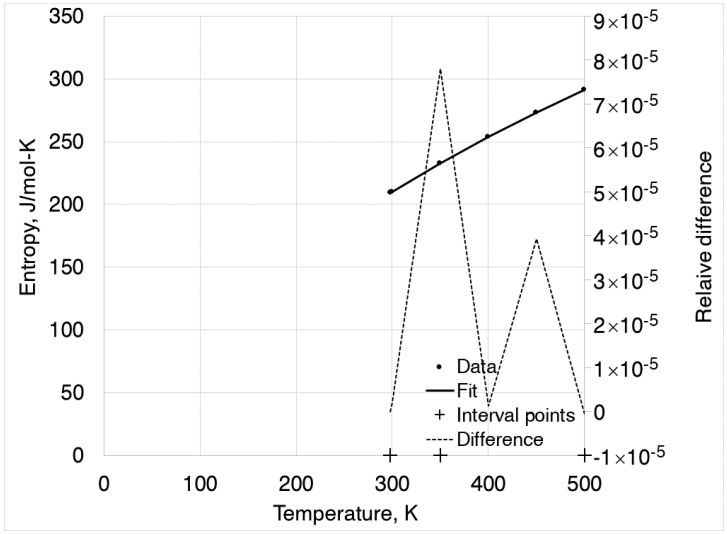


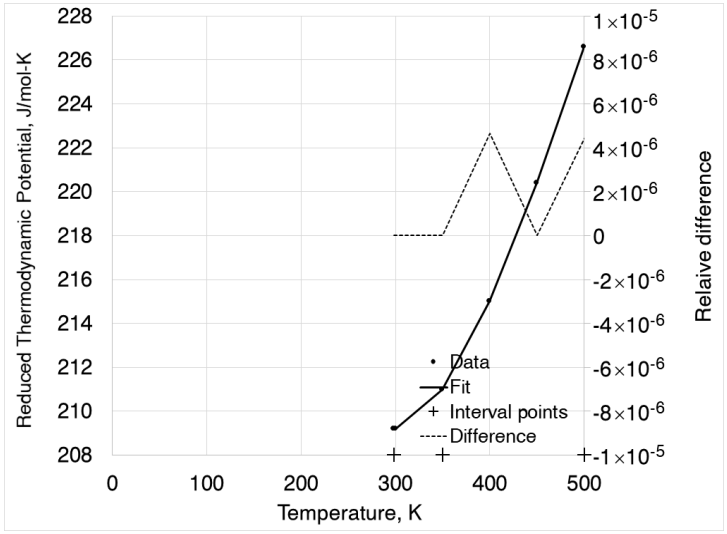
Nitrogen Oxide (N_2O_4) – N2O4(cr,l)

MW = 92.011 g/mol – $\Delta_f H^o$(298.15 K) = -19.564 kJ/mol – S^o(298.15 K) = 209.198 J/mol-K – p^o = 0.1 MPa



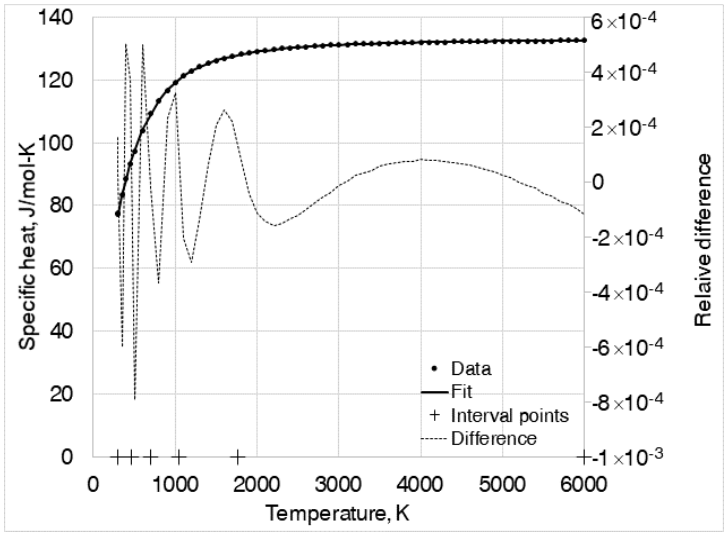


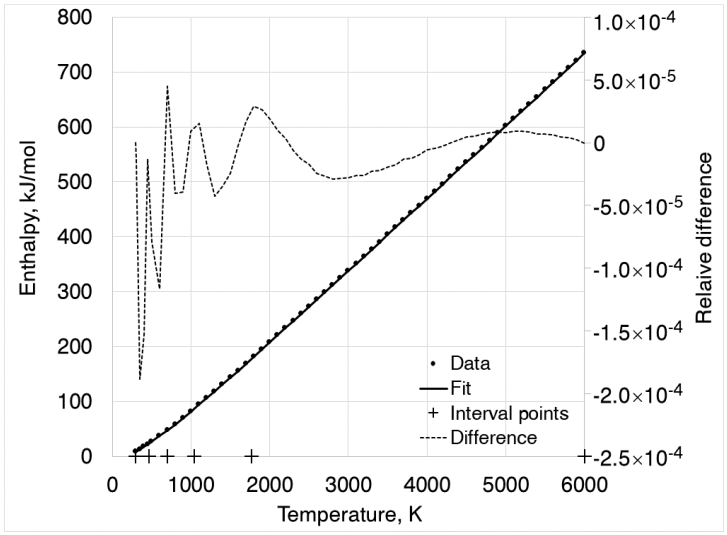


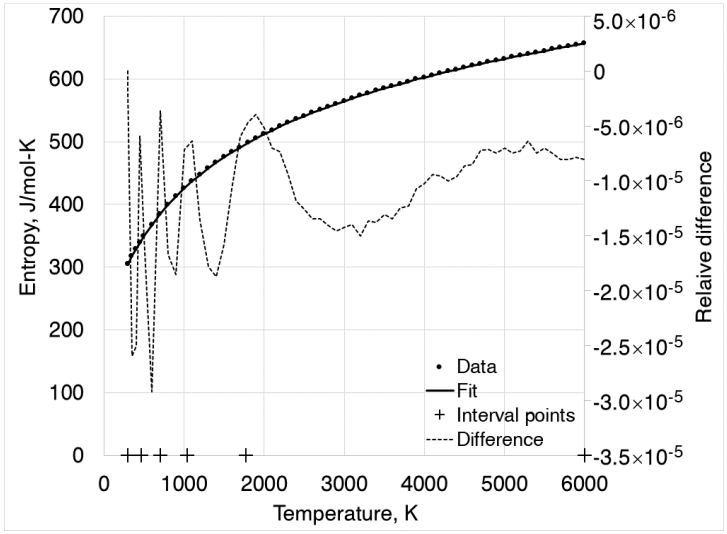


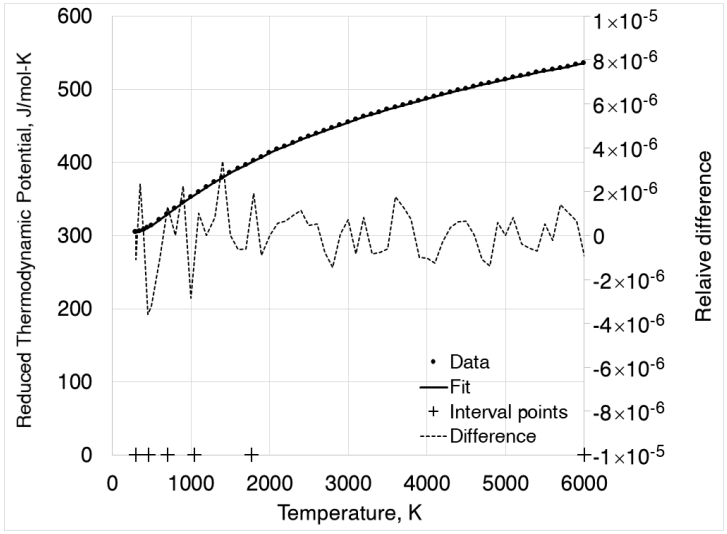
Nitrogen Oxide (N_2O_4) – N2O4(g)

MW = 92.011 g/mol – $\Delta_f H^o$(298.15 K) = 9.079 kJ/mol – S^o(298.15 K) = 304.376 J/mol-K – p^o = 0.1 MPa



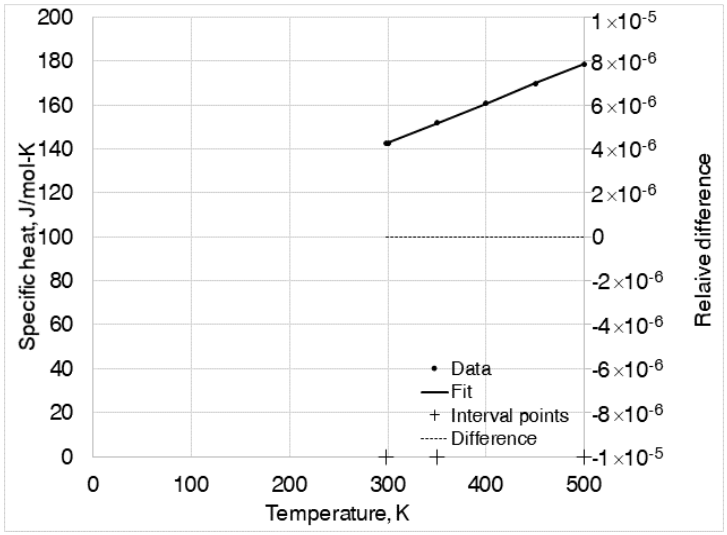


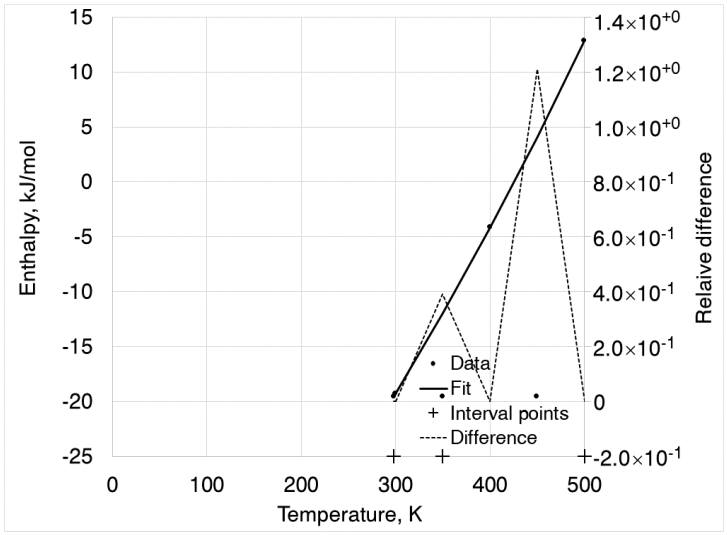


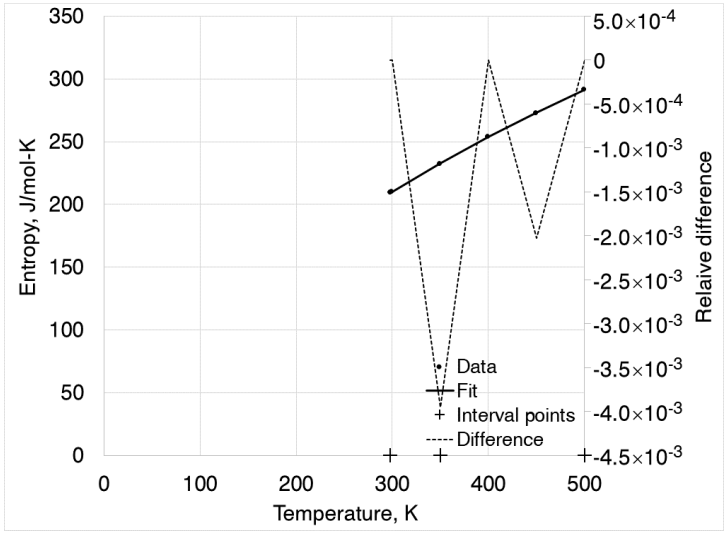


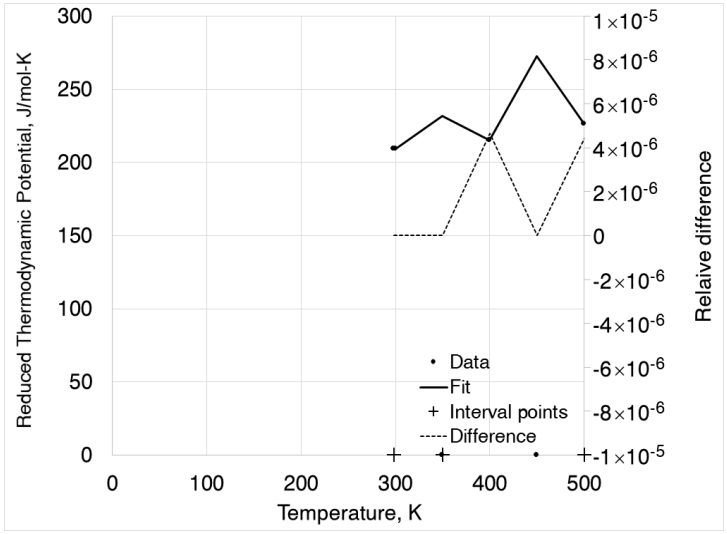
Nitrogen Oxide (N_2O_4) – N2O4(l)

MW = 92.011 g/mol – $\Delta_f H^o$(298.15 K) = -19.564 kJ/mol – S^o(298.15 K) = 209.198 J/mol-K – p^o = 0.1 MPa



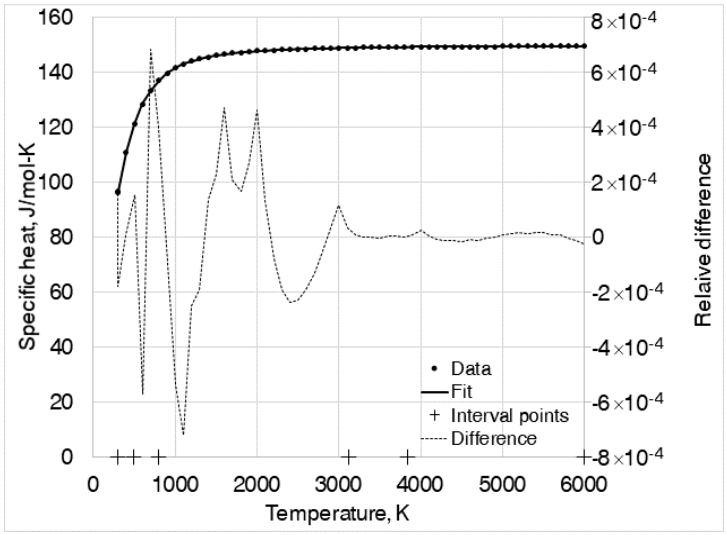


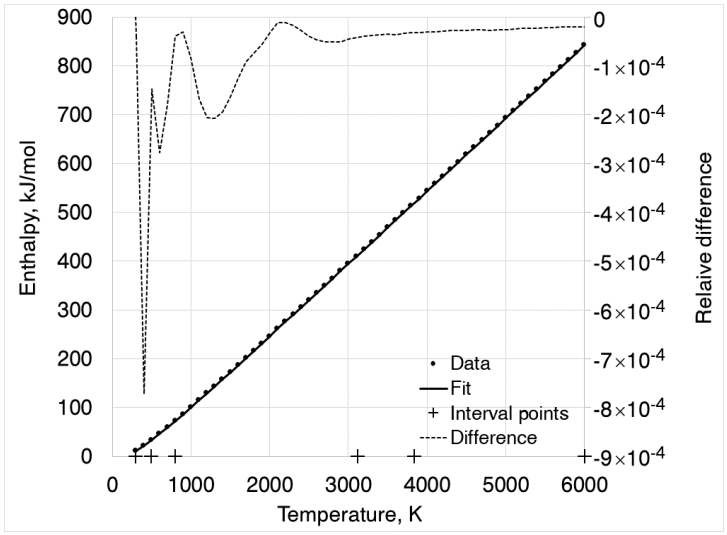


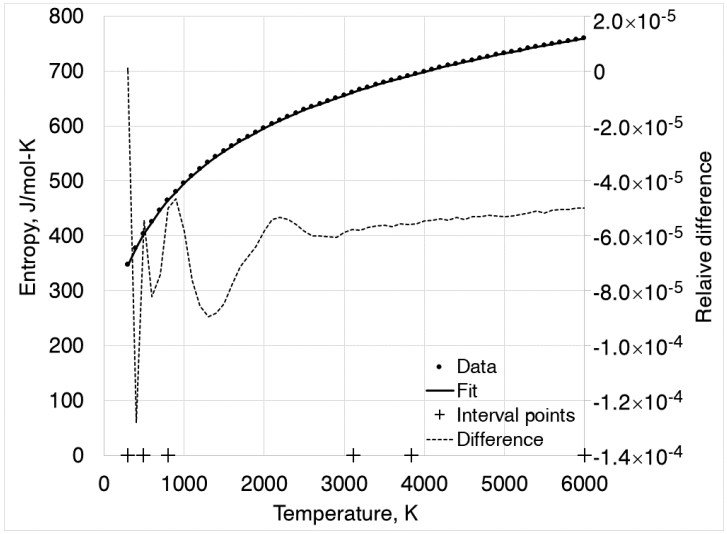


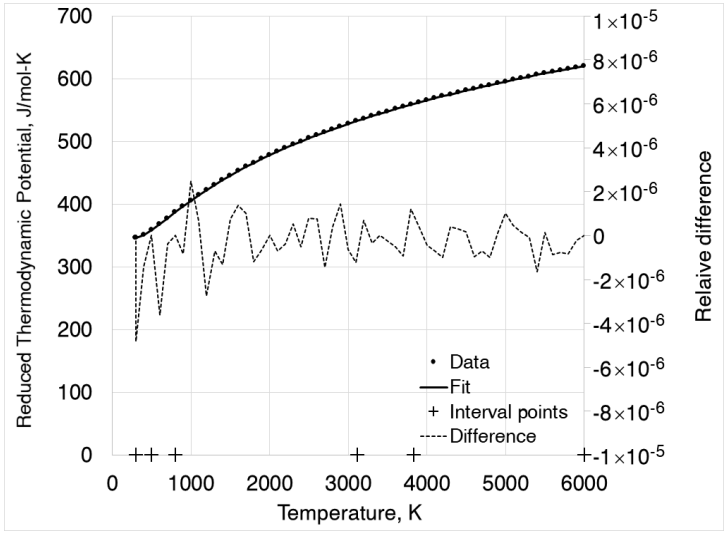
Nitrogen Oxide (N_2O_5) – N2O5(g)

MW = 108.0104 g/mol – $\Delta_f H^o$(298.15 K) = 11.297 kJ/mol – S^o(298.15 K) = 346.548 J/mol-K – p^o = 0.1 MPa



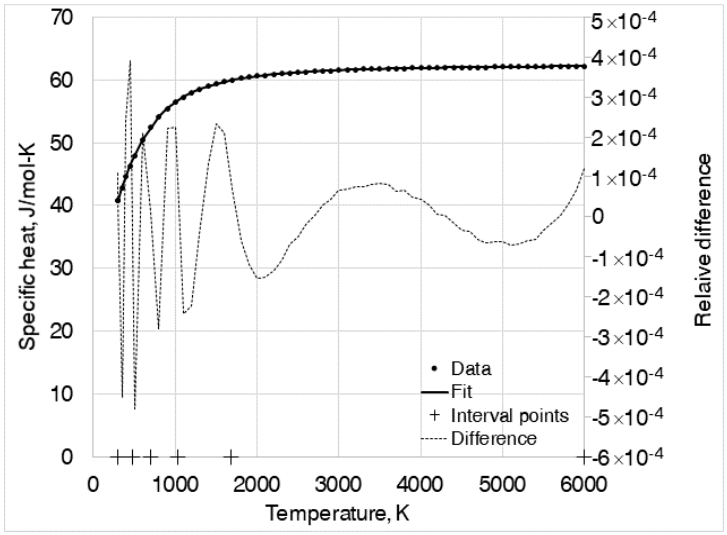


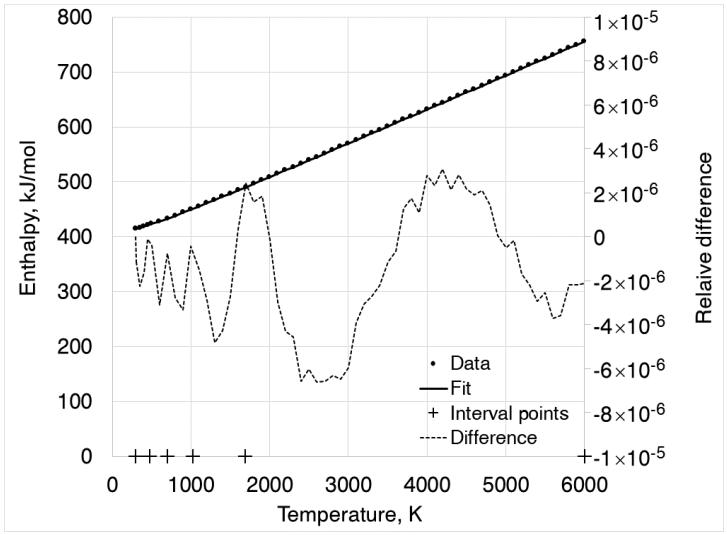


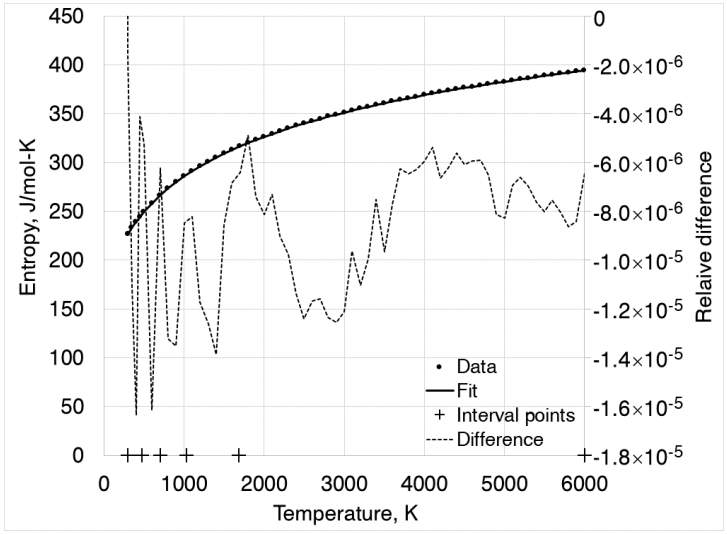


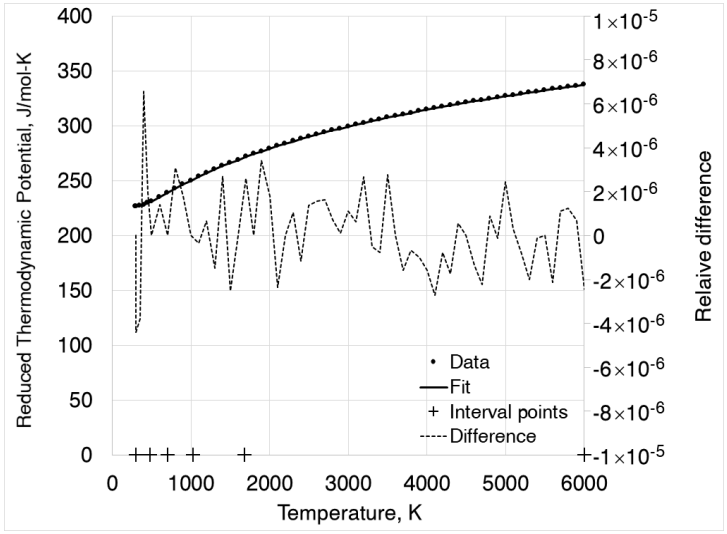
Azide (N_3) – N3(g)

MW = 42.0201 g/mol – $\Delta_f H^o$(298.15 K) = 414.216 kJ/mol – S^o(298.15 K) = 226.469 J/mol-K – p^o = 0.1 MPa



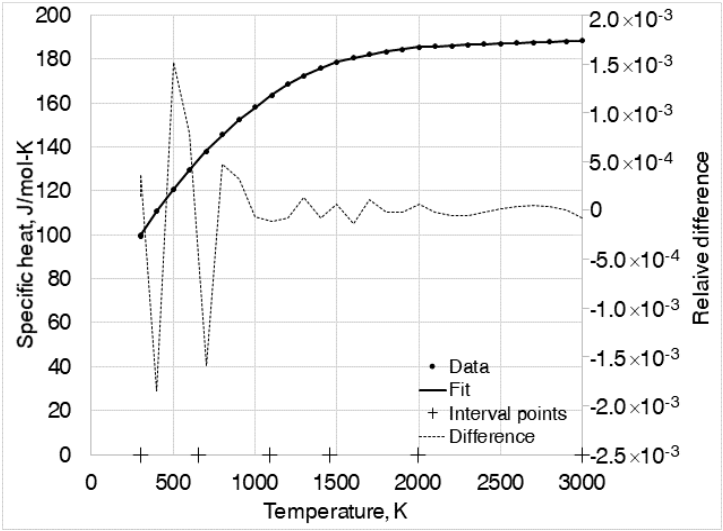


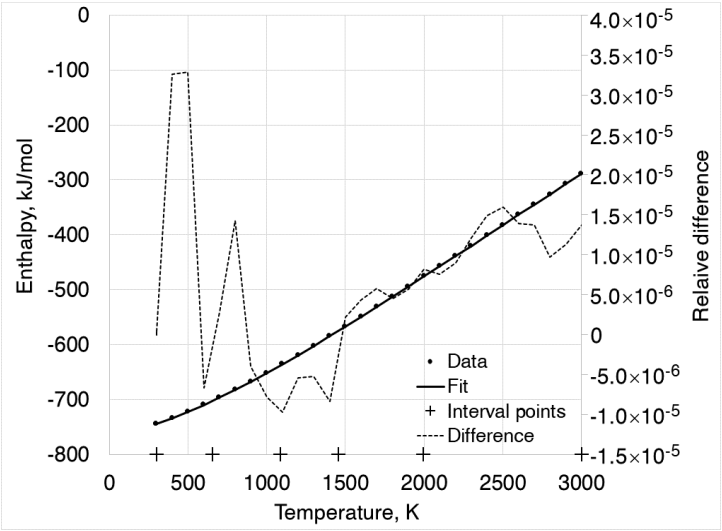


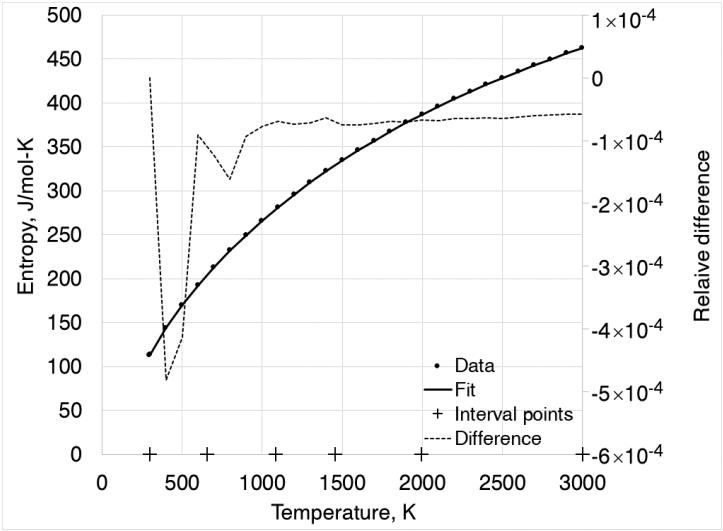


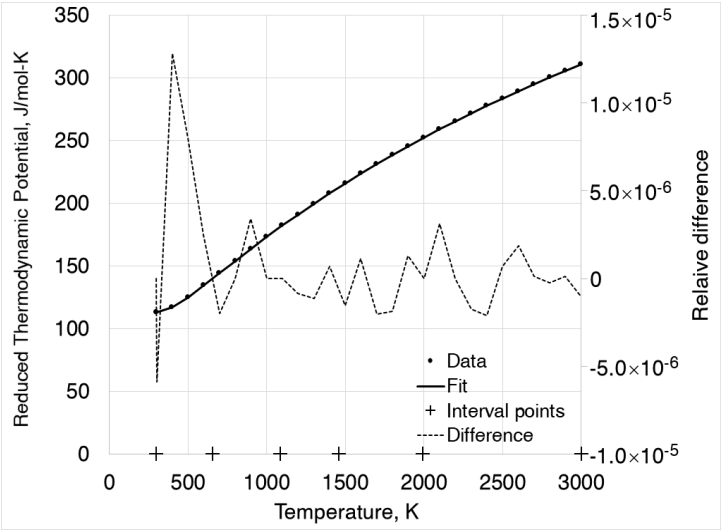
Silicon Nitride, Alpha (Si_3N_4) – N4Si3(cr)

MW = 140.2833 g/mol – $\Delta_f H^o$(298.15 K) = -744.752 kJ/mol – S^o(298.15 K) = 112.967 J/mol-K – p^o = 0.1 MPa



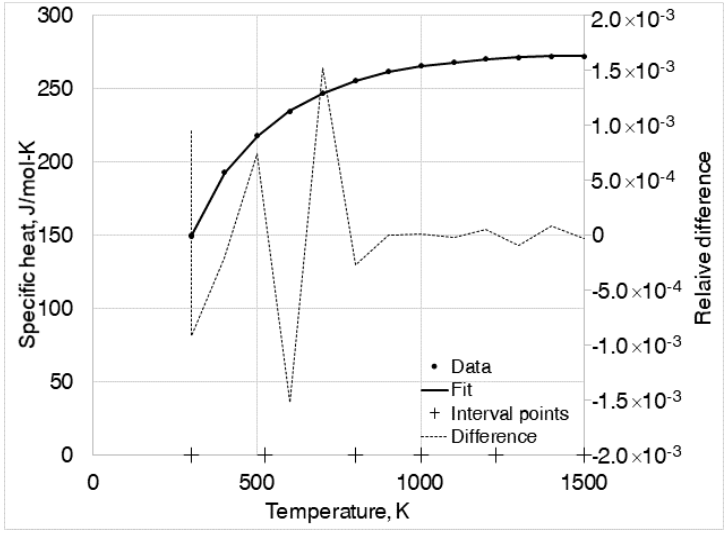


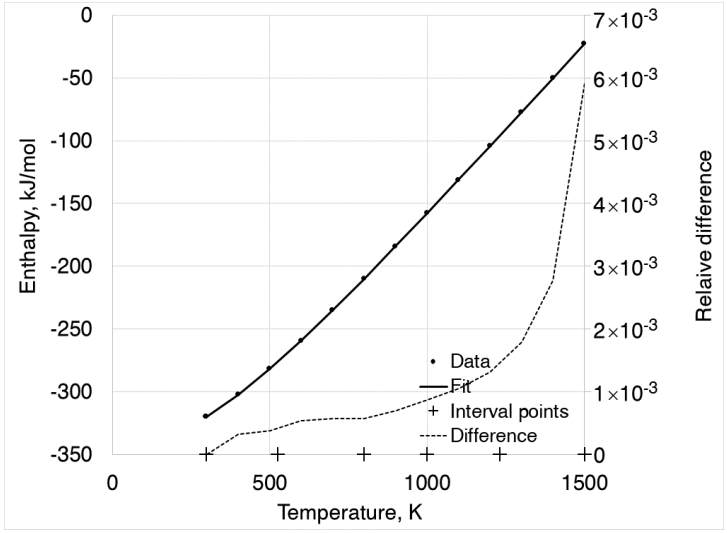


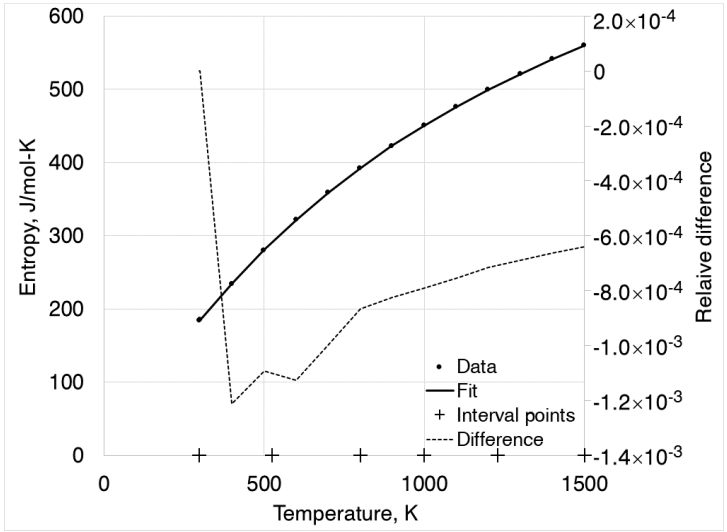


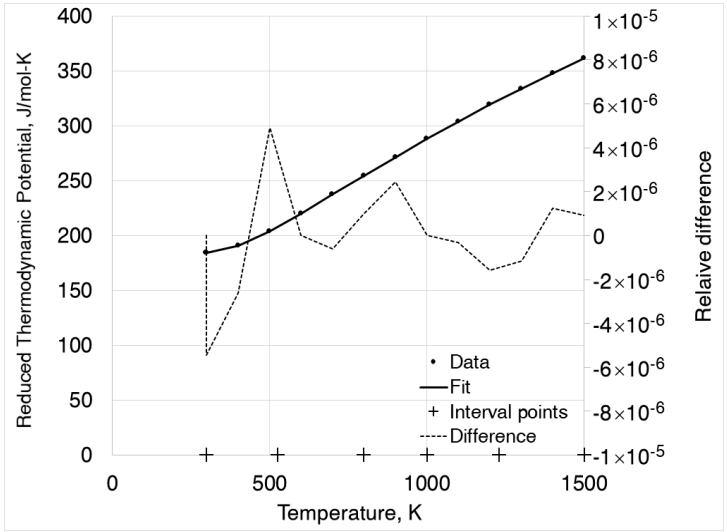
Phosphorus Nitride (P_3N_5) – N5P3(cr)

MW = 162.9548 g/mol – $\Delta_f H^o$(298.15 K) = -320.155 kJ/mol – S^o(298.15 K) = 184.096 J/mol-K – p^o = 0.1 MPa



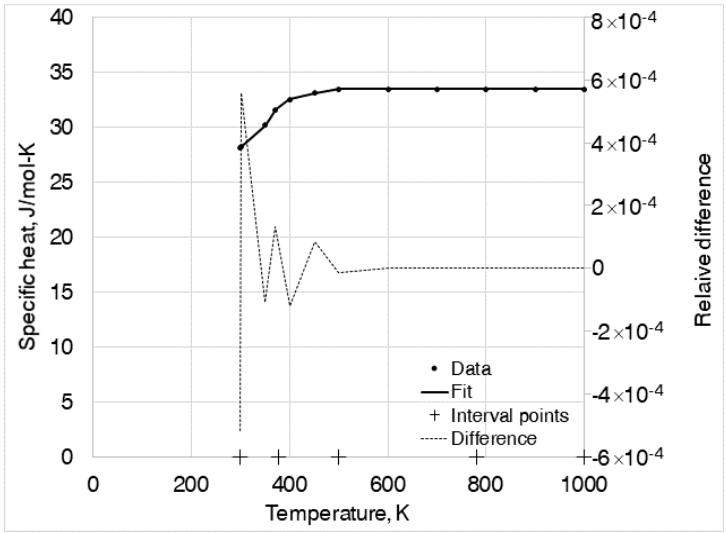


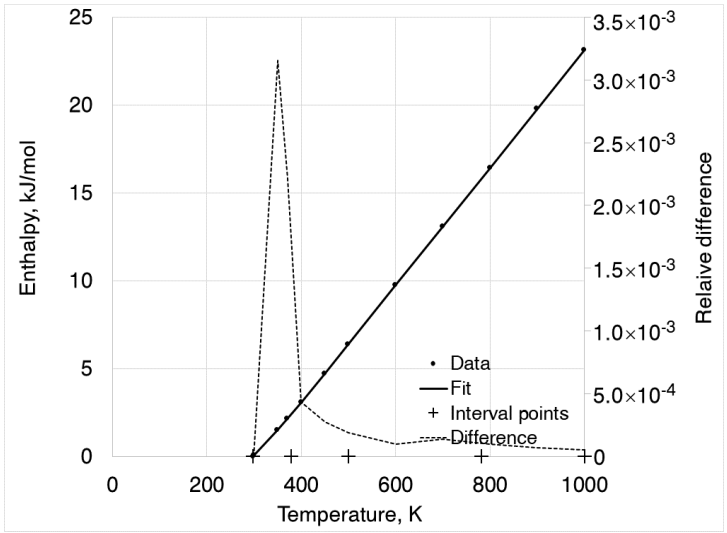


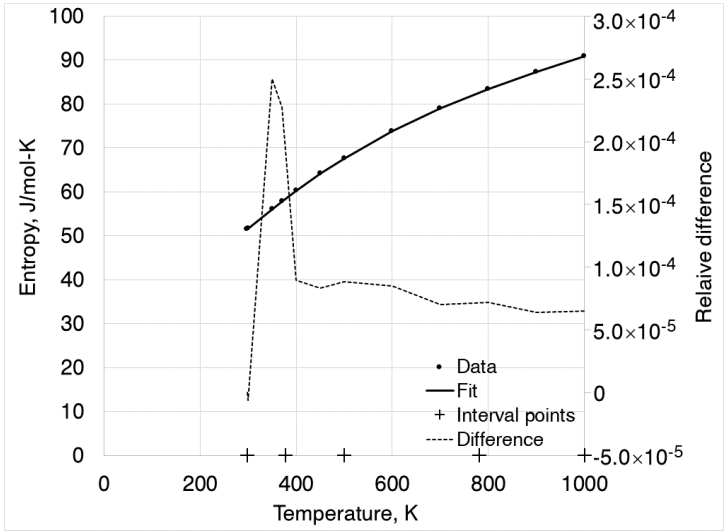


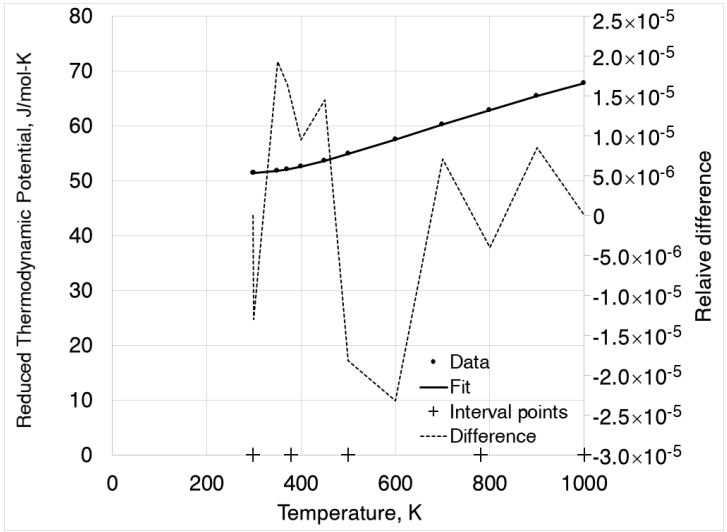
Sodium (Na) – Na1(cr)

MW = 22.98977 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 51.455 J/mol-K – p^o = 0.1 MPa



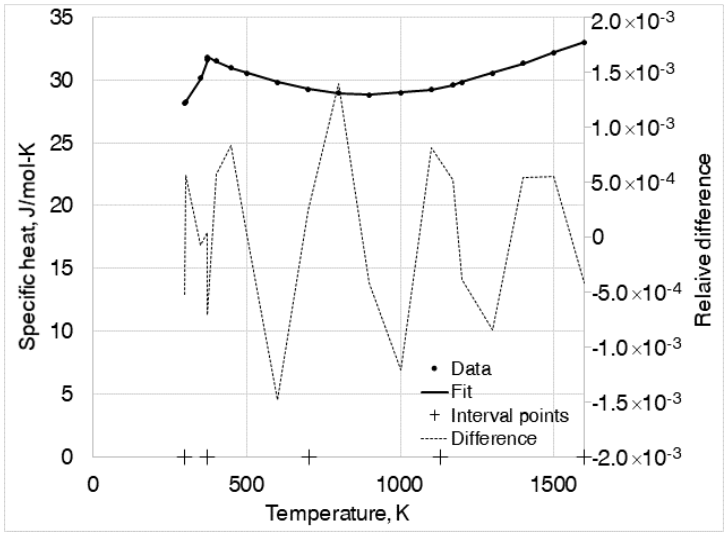


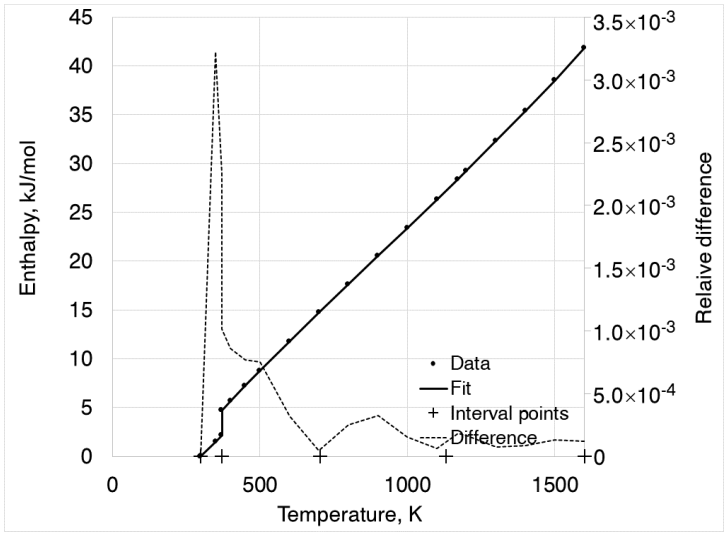


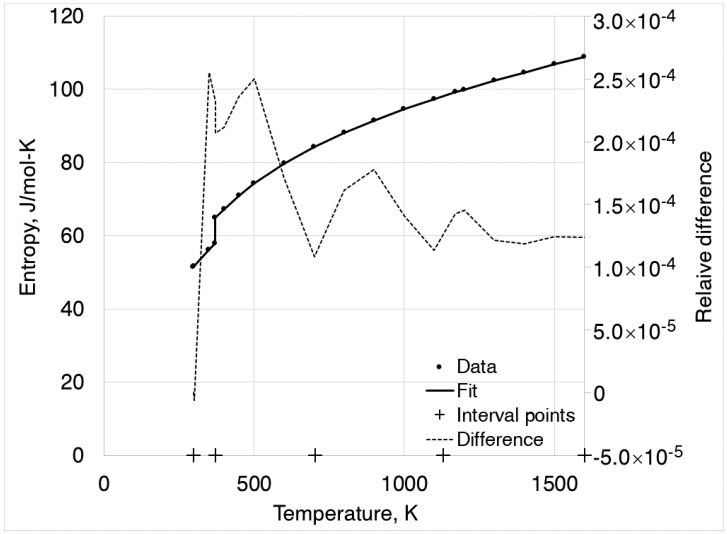


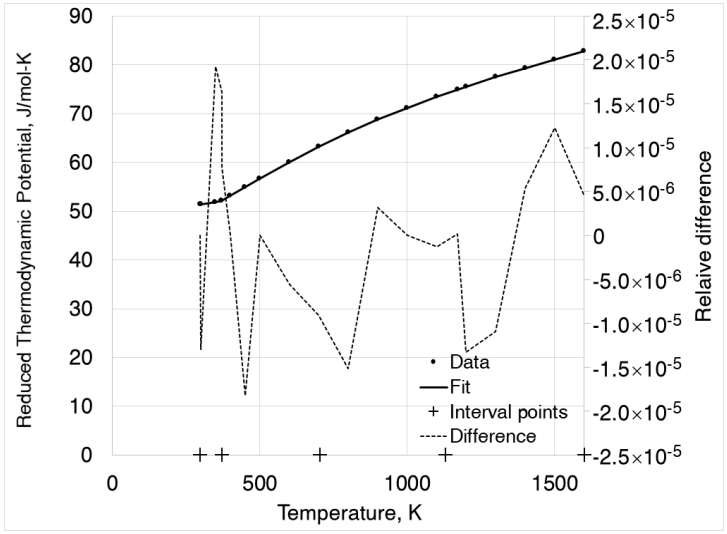
Sodium (Na) – Na1(cr,l)

MW = 22.98977 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 51.455 J/mol-K – p^o = 0.1 MPa



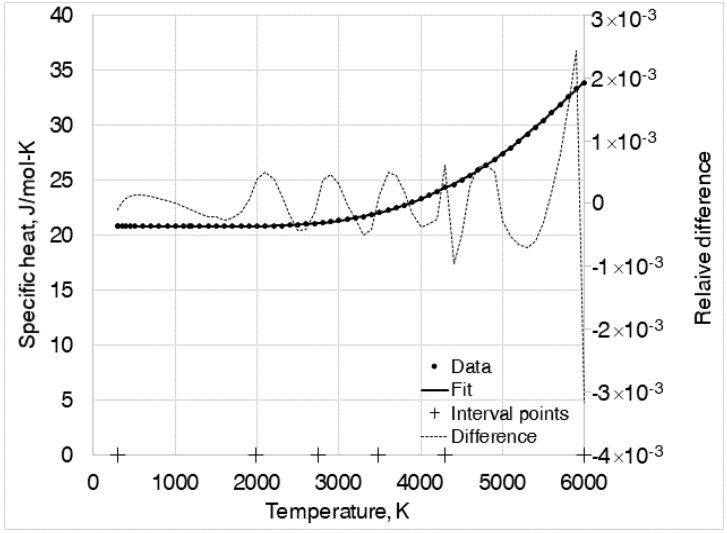


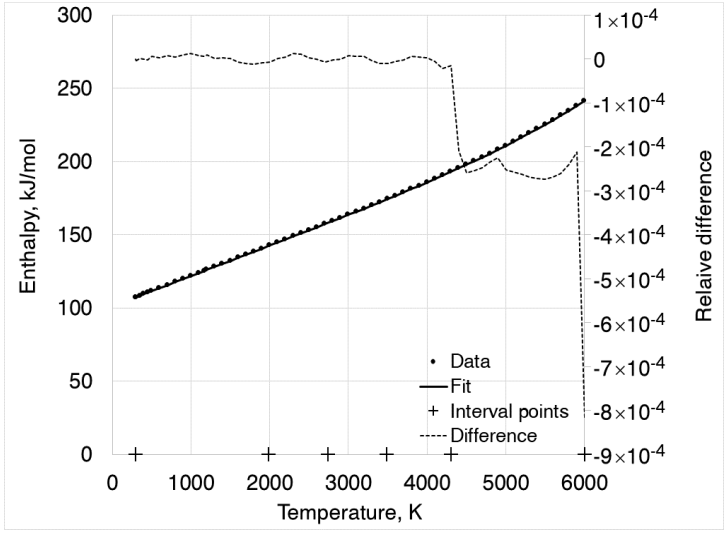


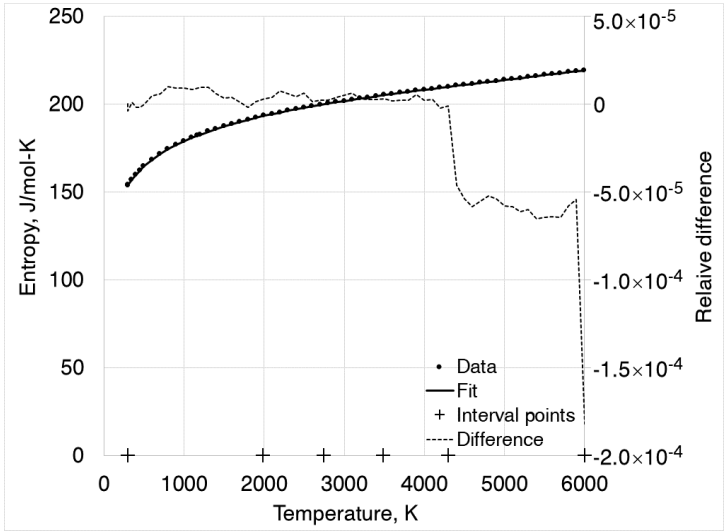


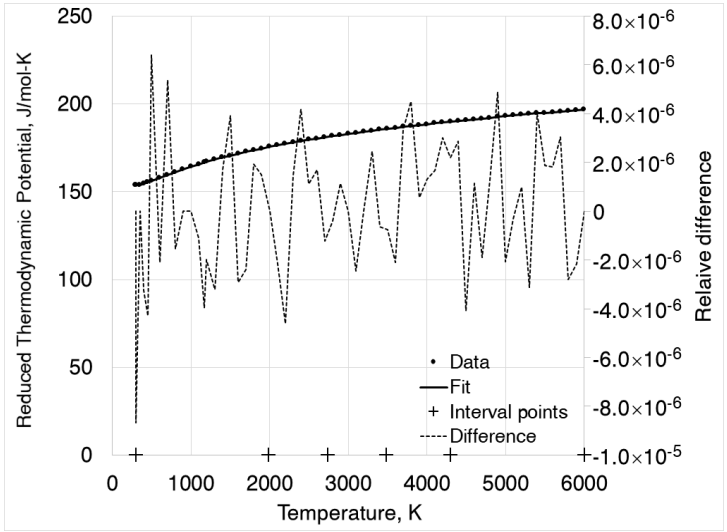
Sodium (Na) – Na1(g)

MW = 22.98977 g/mol – $\Delta_f H^o$(298.15 K) = 107.3 kJ/mol – S^o(298.15 K) = 153.667 J/mol-K – p^o = 0.1 MPa



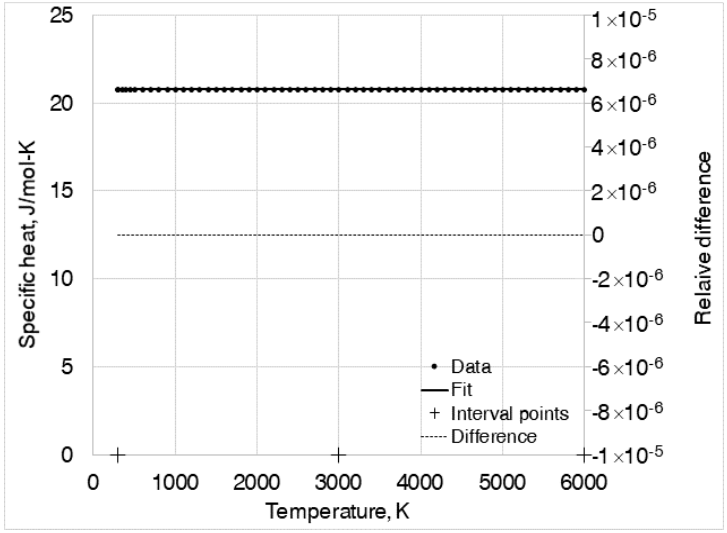


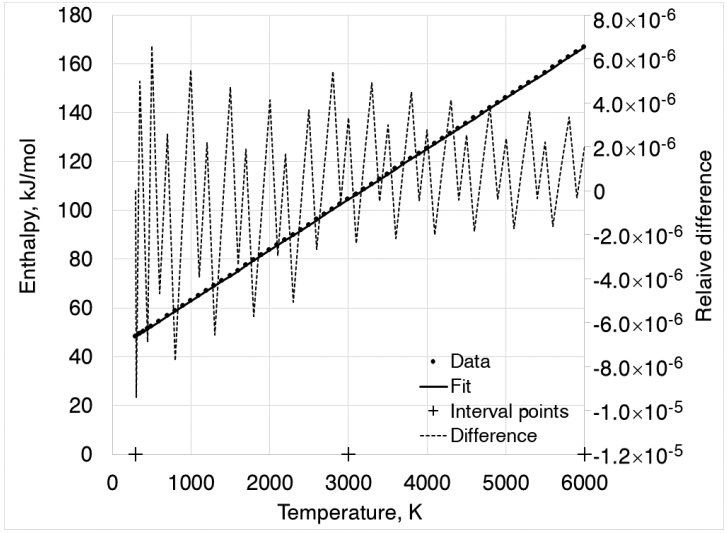


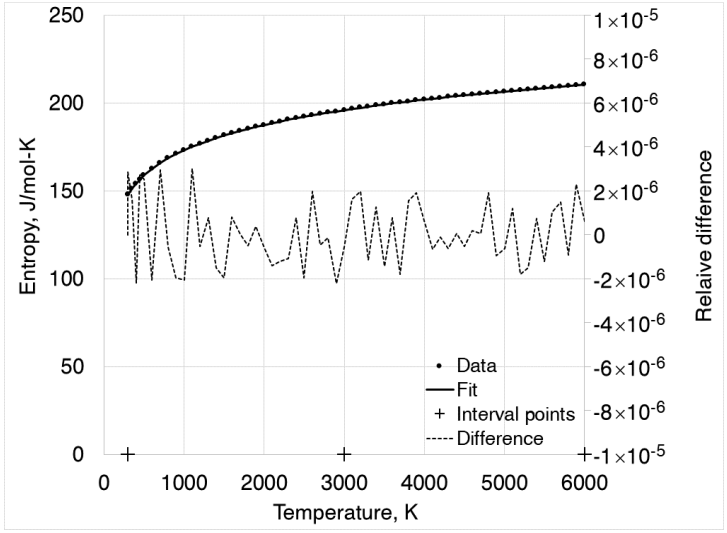


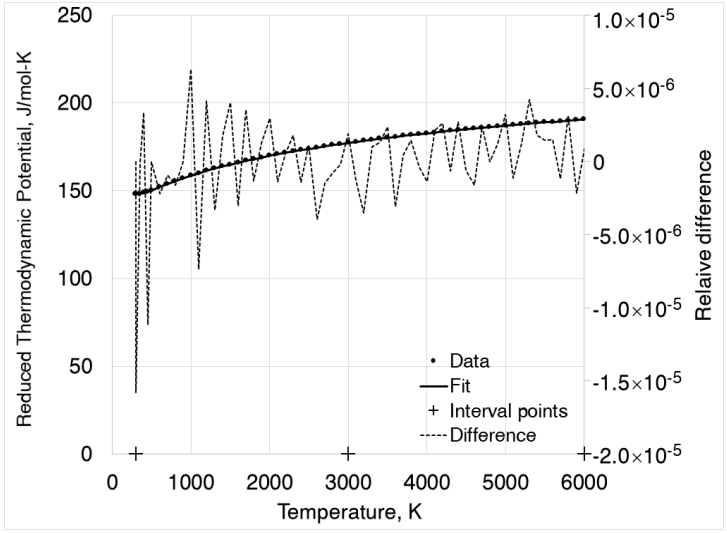
Sodium, Ion (Na^-) – Na1-(g)

MW = 22.99032 g/mol – $\Delta_f H^o$(298.15 K) = 48.236 kJ/mol – S^o(298.15 K) = 147.954 J/mol-K – p^o = 0.1 MPa



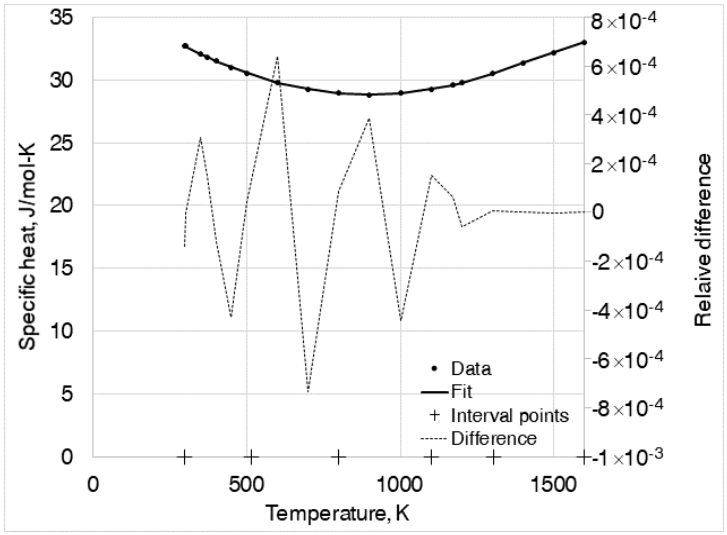


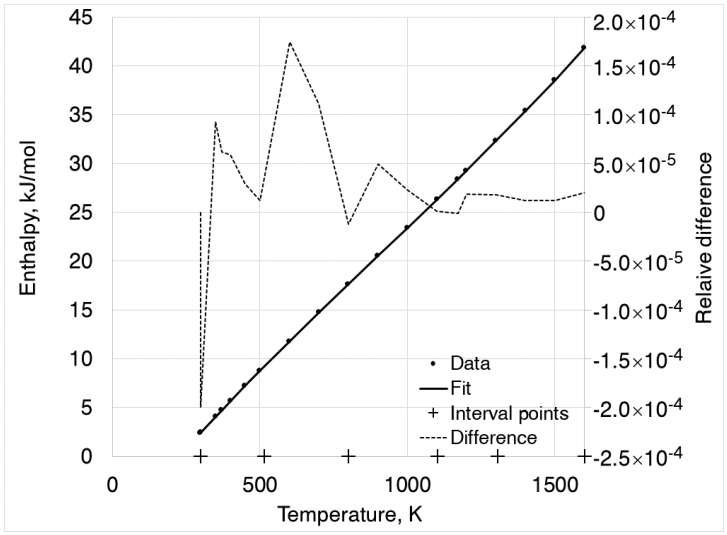


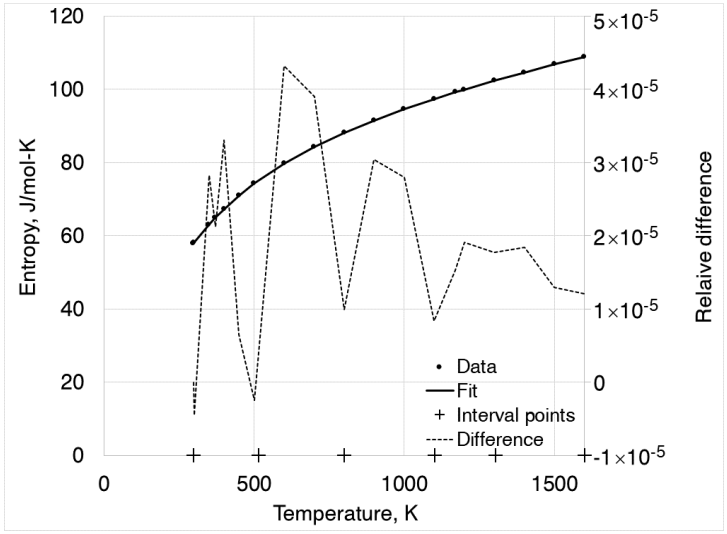


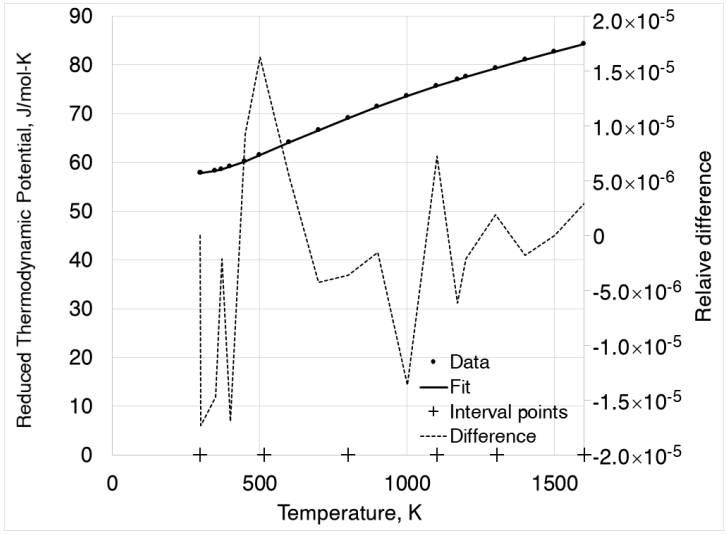
Sodium (Na) – Na1(l)

MW = 22.98977 g/mol – $\Delta_f H^o$(298.15 K) = 2.406 kJ/mol – S^o(298.15 K) = 57.862 J/mol-K – p^o = 0.1 MPa



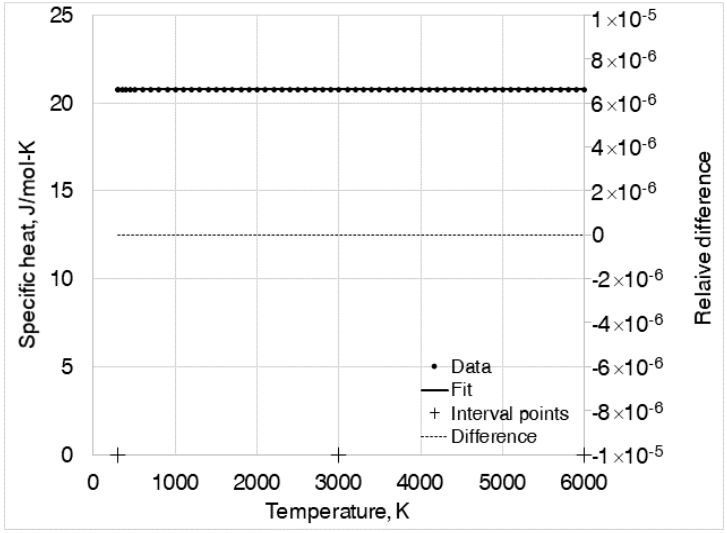


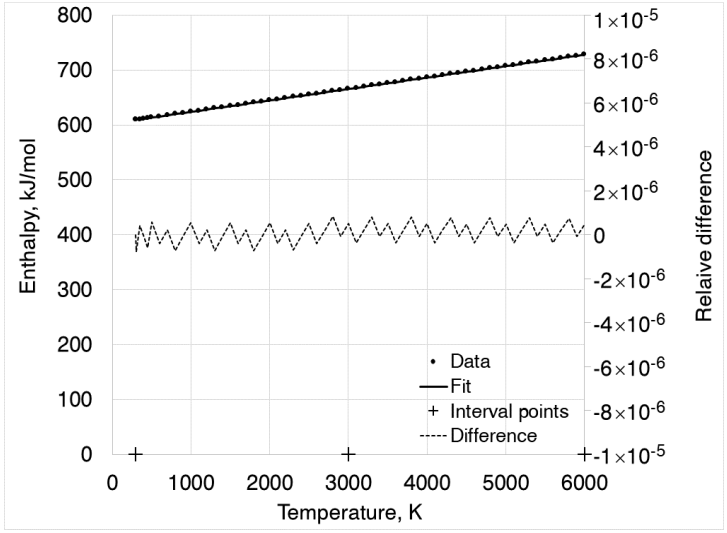


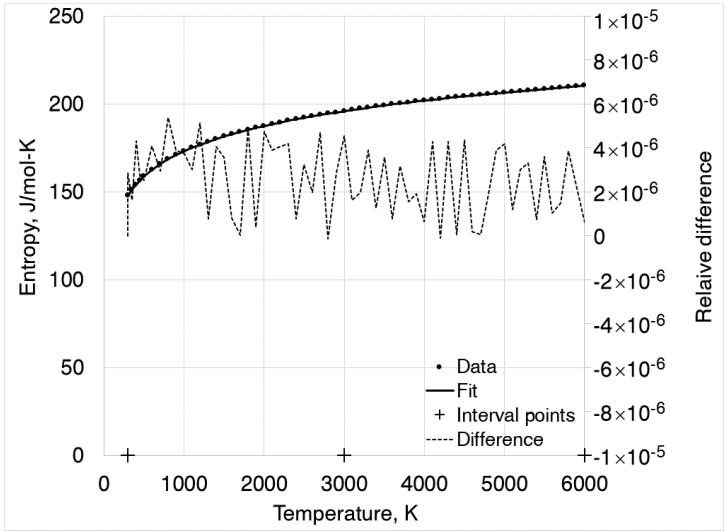


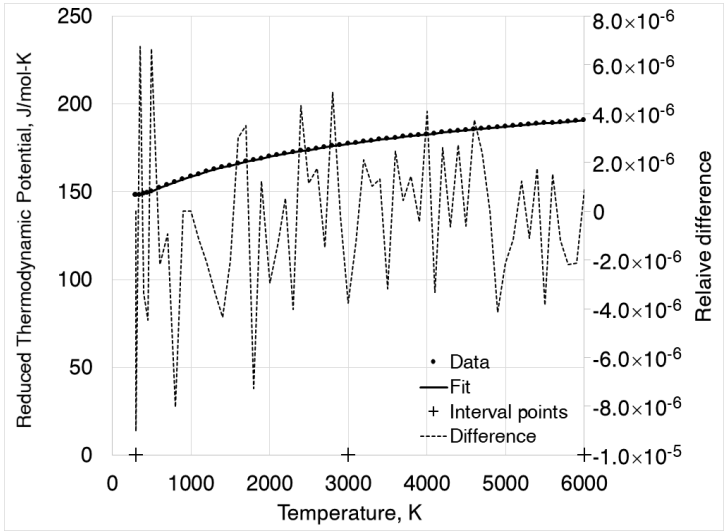
Sodium, Ion (Na^+) – Na1+(g)

MW = 22.98922 g/mol – $\Delta_f H^o$(298.15 K) = 609.343 kJ/mol – S^o(298.15 K) = 147.953 J/mol-K – p^o = 0.1 MPa



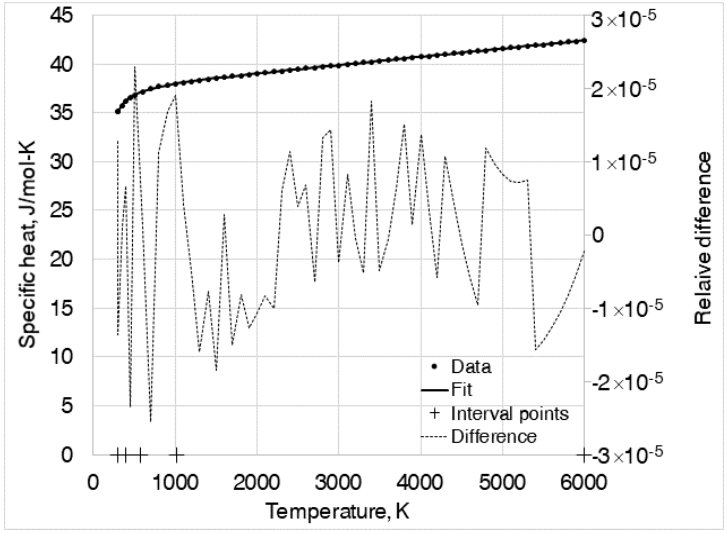


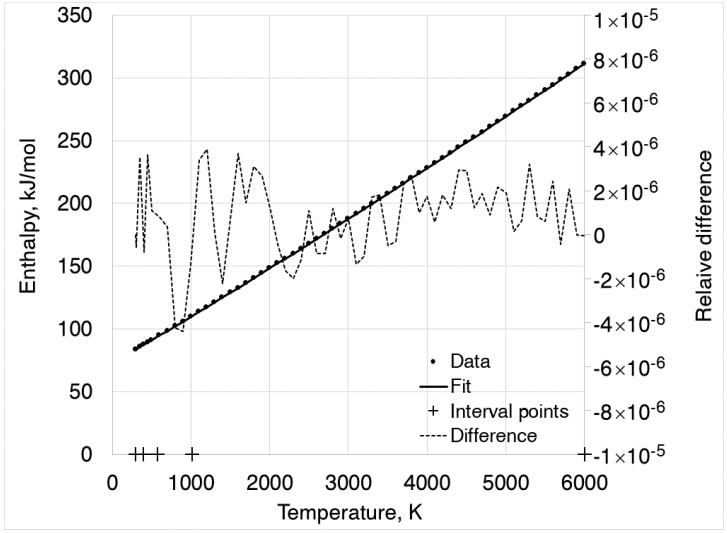


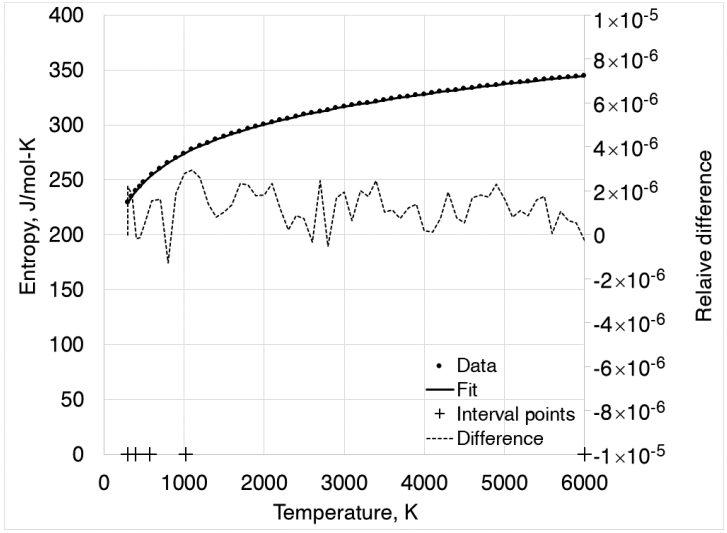


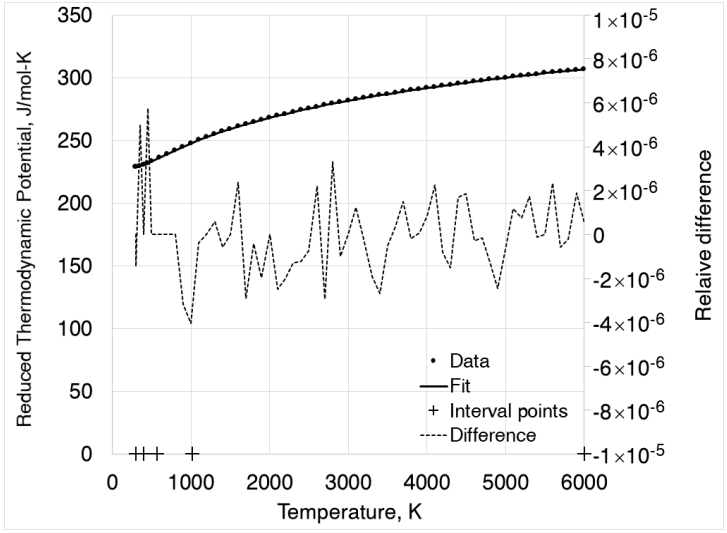
Sodium Oxide (NaO) – Na1O1(g)

MW = 38.9892 g/mol – $\Delta_f H^o$(298.15 K) = 83.68 kJ/mol – S^o(298.15 K) = 229.133 J/mol-K – p^o = 0.1 MPa



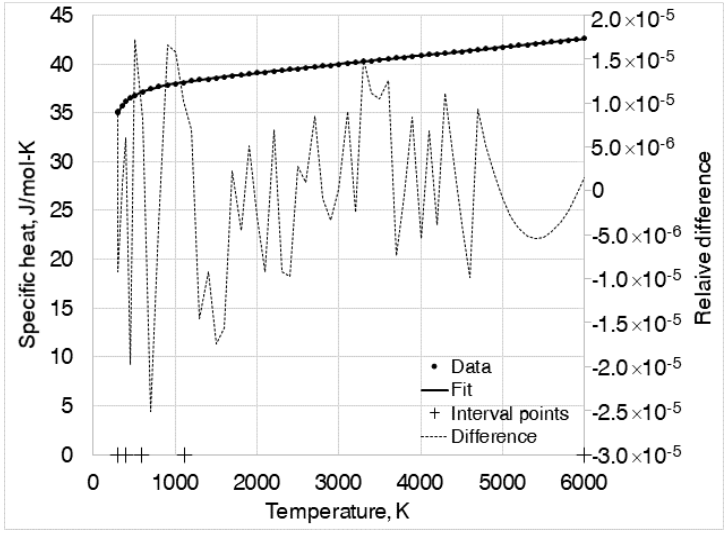


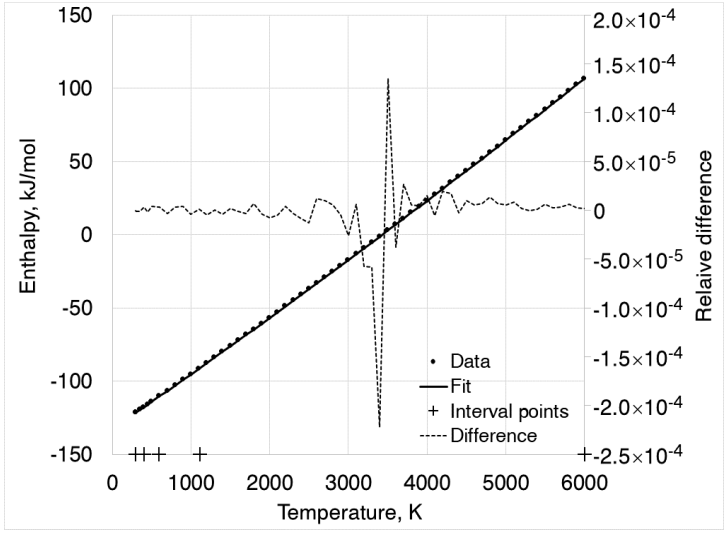


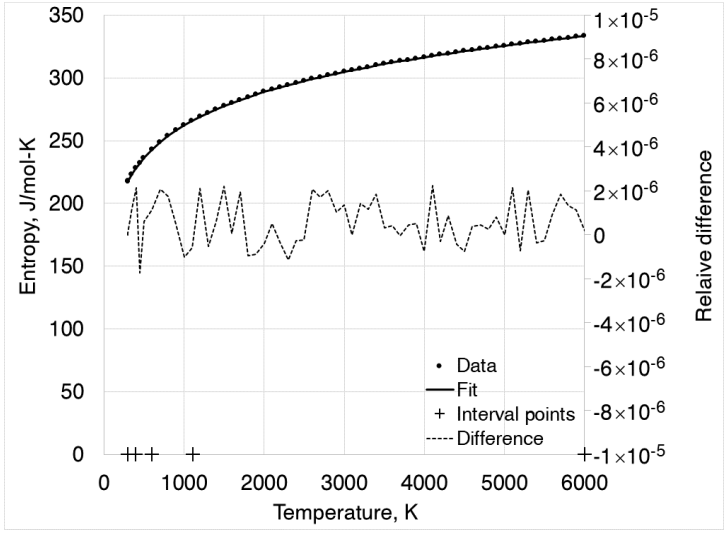


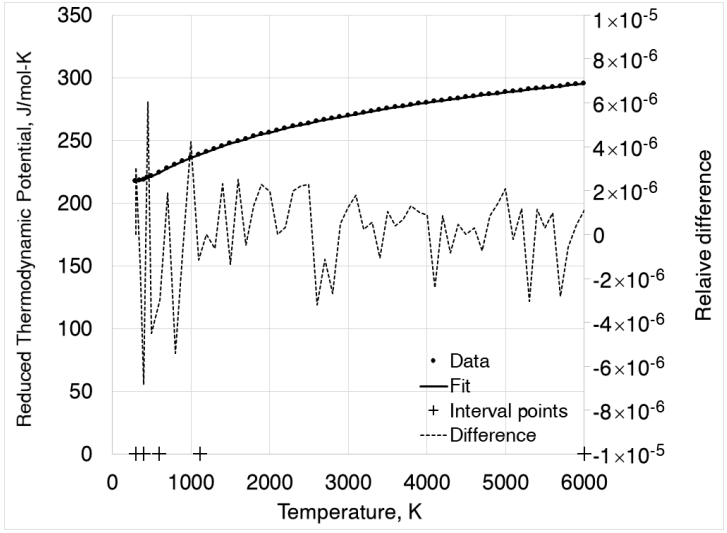
Sodium Oxide, Ion (NaO^-) – Na1O1-(g)

MW = 38.9897 g/mol – $\Delta_f H^o$(298.15 K) = -121.336 kJ/mol – S^o(298.15 K) = 217.487 J/mol-K – p^o = 0.1 MPa



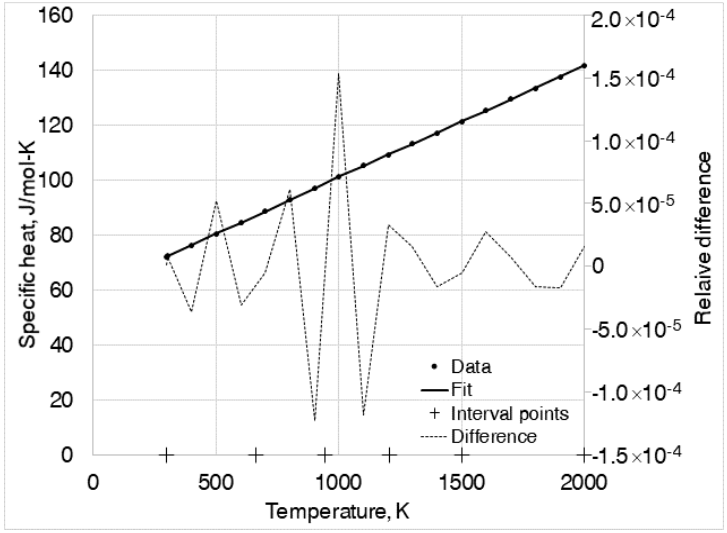


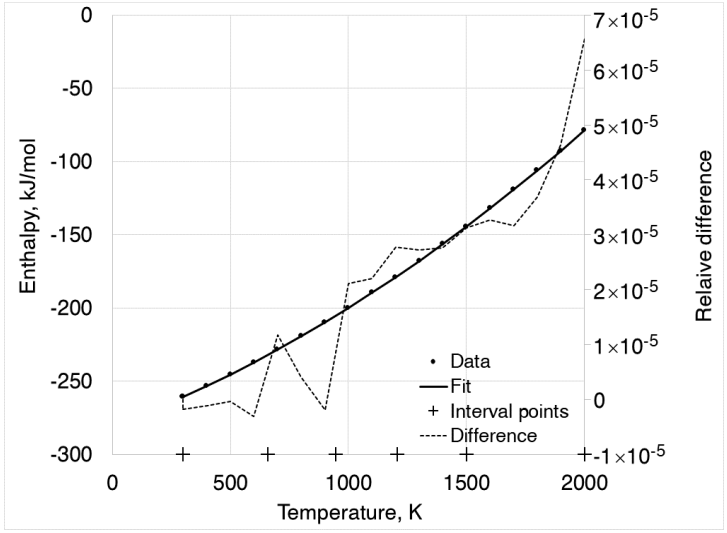


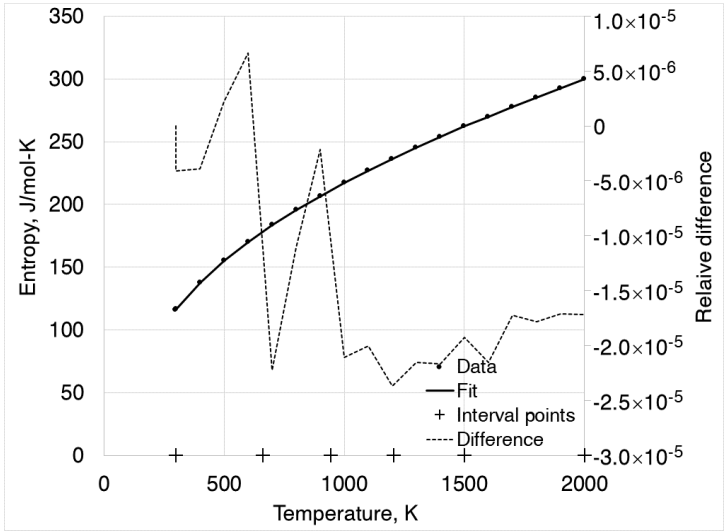


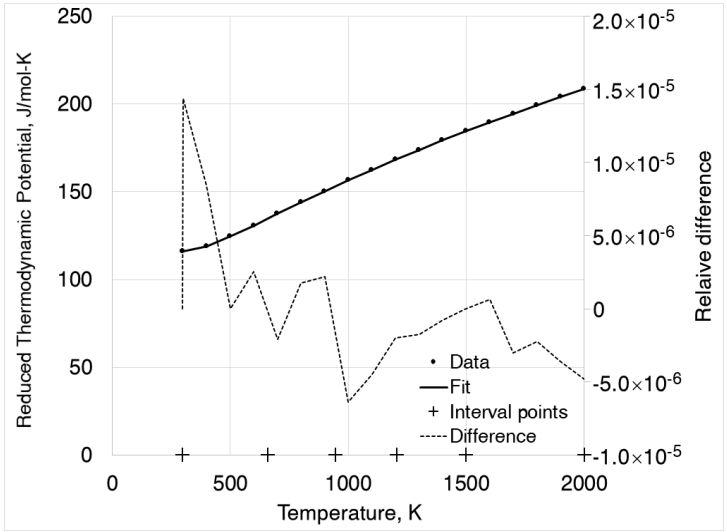
Sodium Superoxide (NaO_2) – Na1O2(cr)

MW = 54.9886 g/mol – $\Delta_f H^o$(298.15 K) = -260.663 kJ/mol – S^o(298.15 K) = 115.897 J/mol-K – p^o = 0.1 MPa



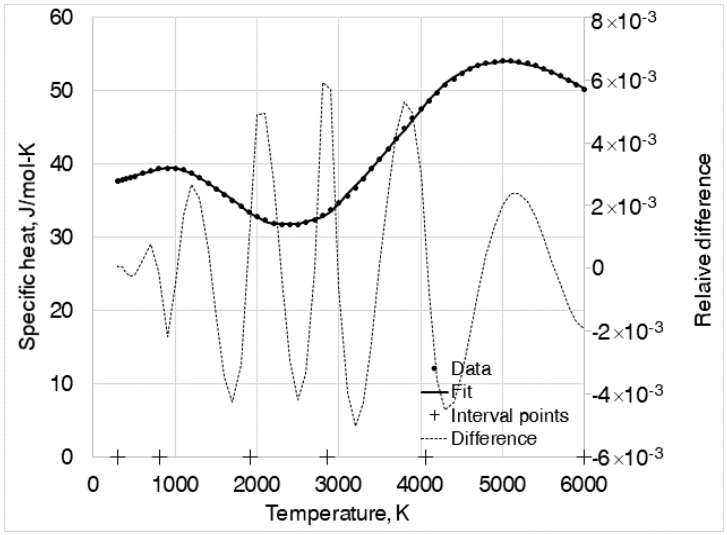


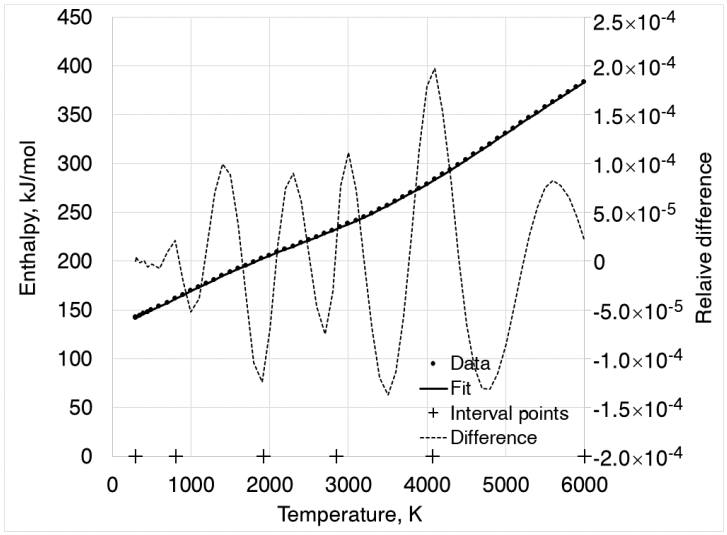


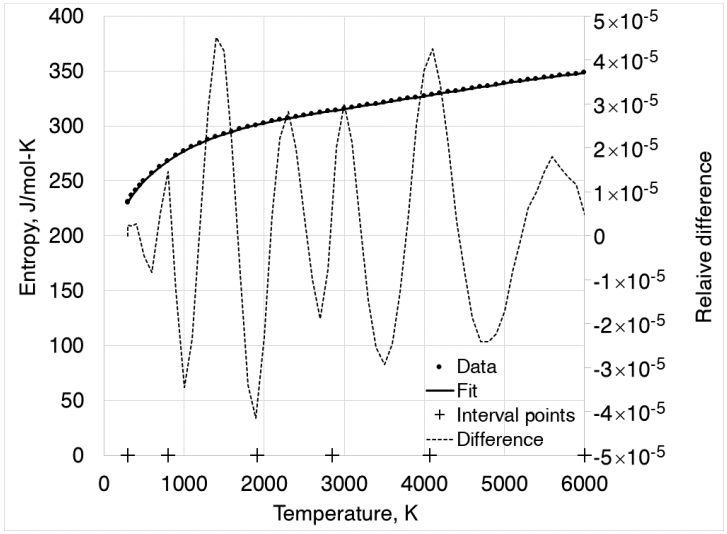


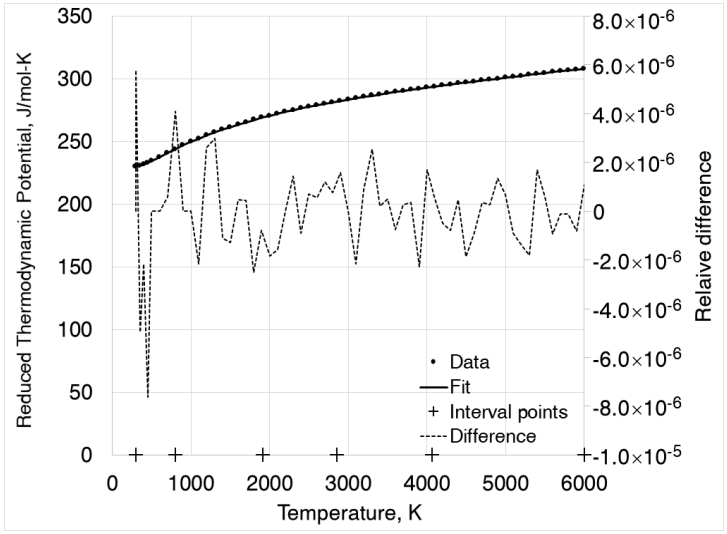
Sodium (Na_2) – Na2(g)

MW = 45.97954 g/mol – $\Delta_f H^o$(298.15 K) = 142.07 kJ/mol – S^o(298.15 K) = 230.243 J/mol-K – p^o = 0.1 MPa



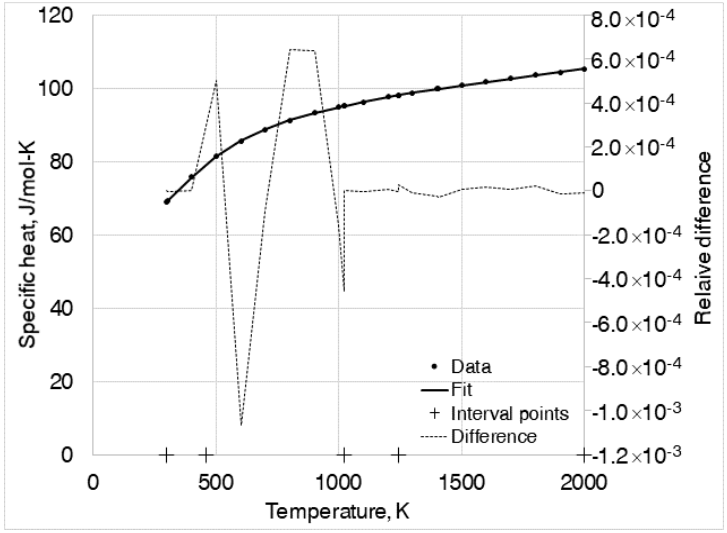


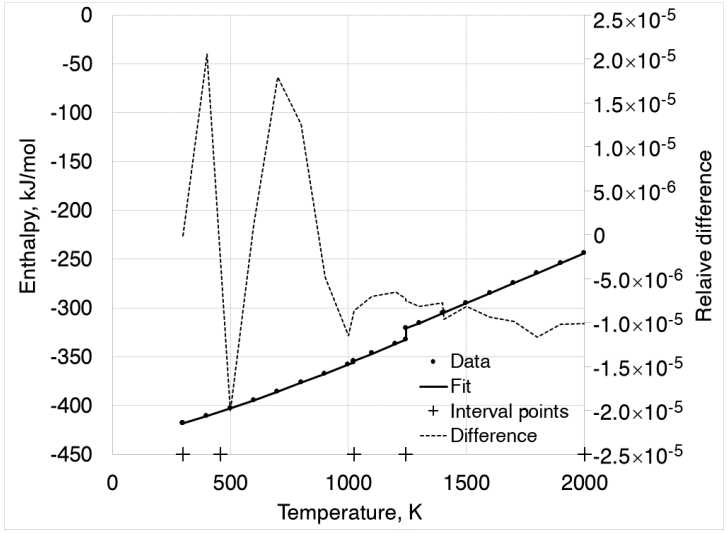


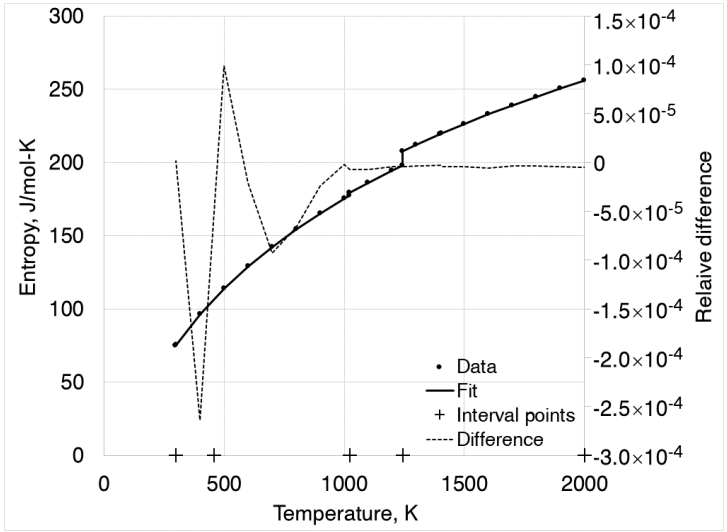


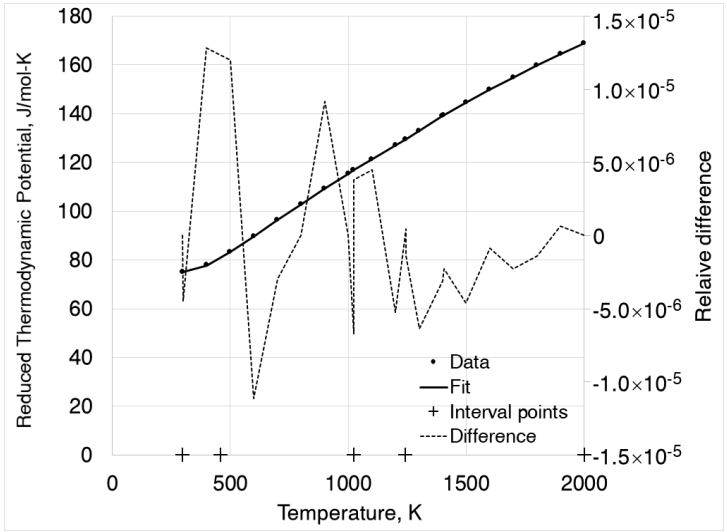
Sodium Oxide (Na_2O) – Na2O1(cr)

MW = 61.9789 g/mol – $\Delta_f H^o$(298.15 K) = -417.982 kJ/mol – S^o(298.15 K) = 75.042 J/mol-K – p^o = 0.1 MPa



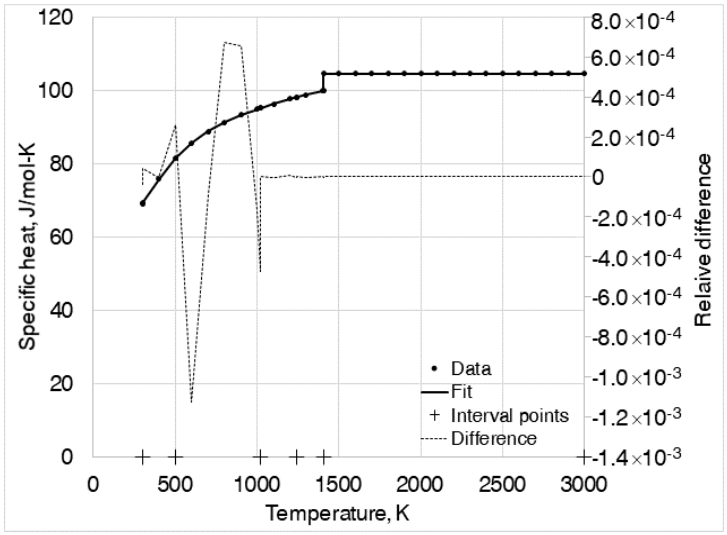


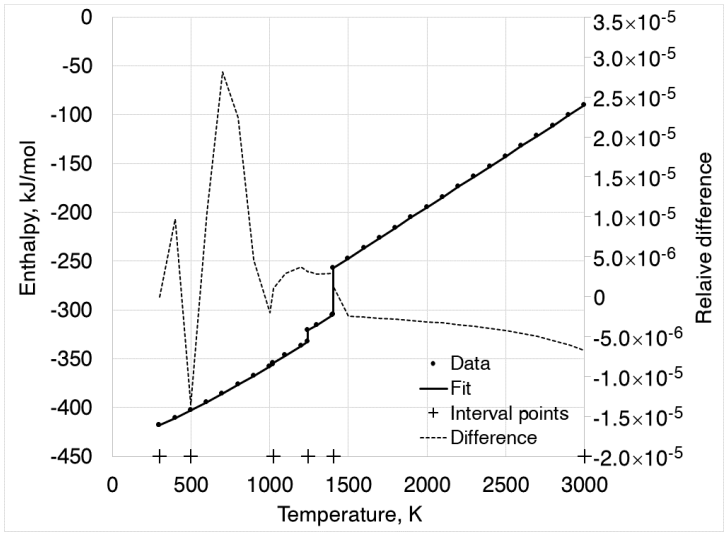


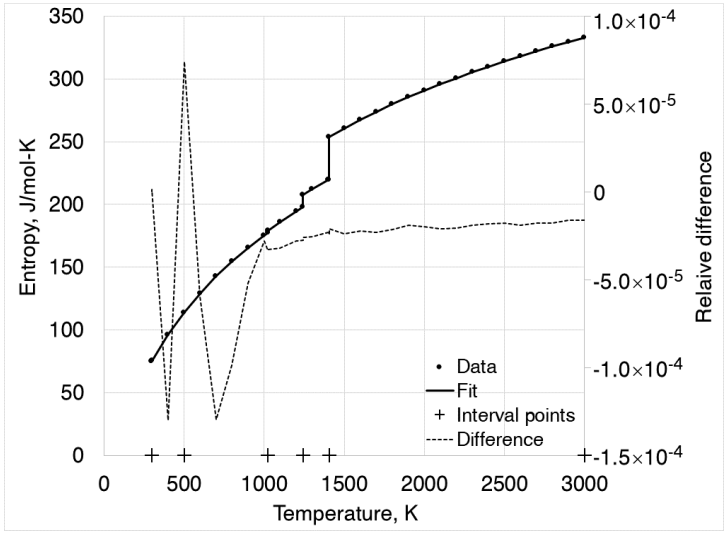


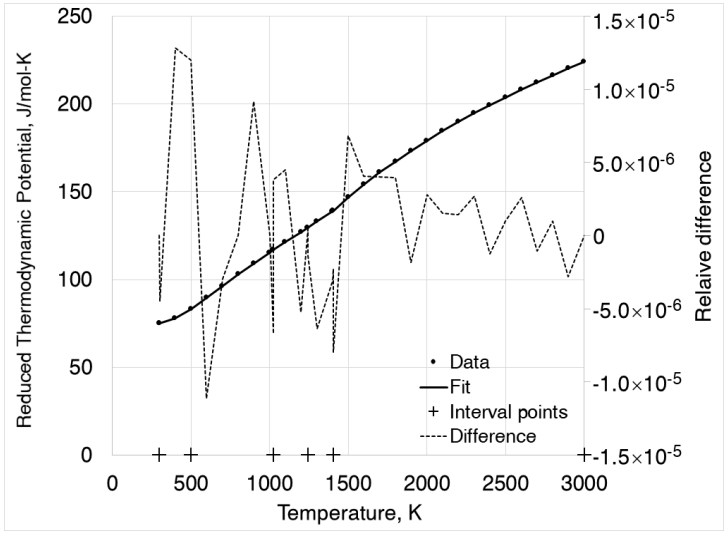
Sodium Oxide (Na_2O) – Na2O1(cr,l)

MW = 61.9789 g/mol – $\Delta_f H^o$(298.15 K) = -417.982 kJ/mol – S^o(298.15 K) = 75.042 J/mol-K – p^o = 0.1 MPa



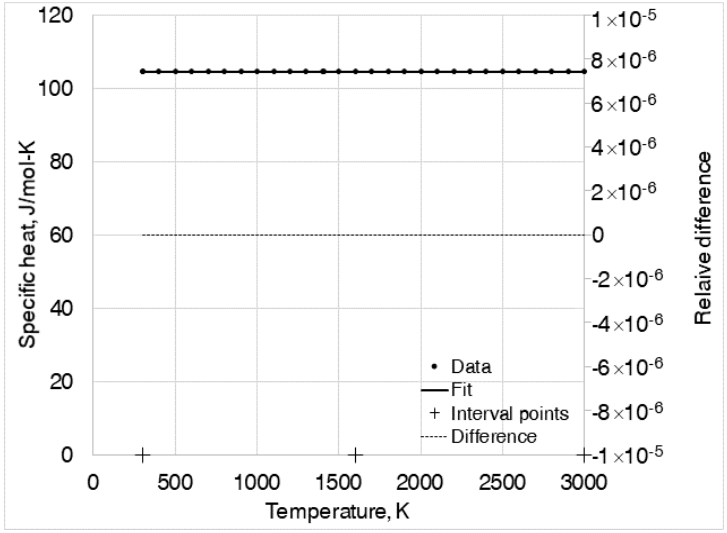


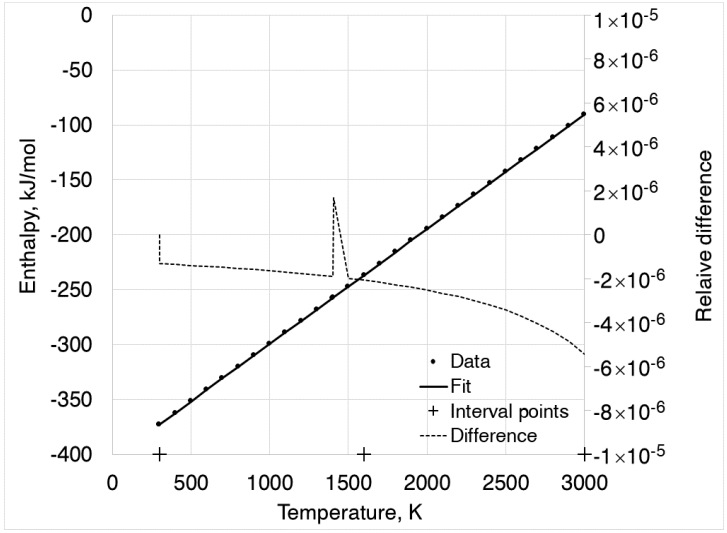


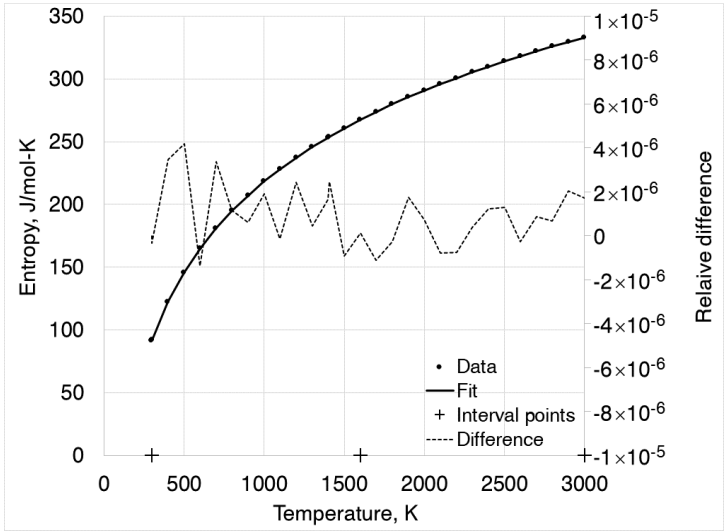


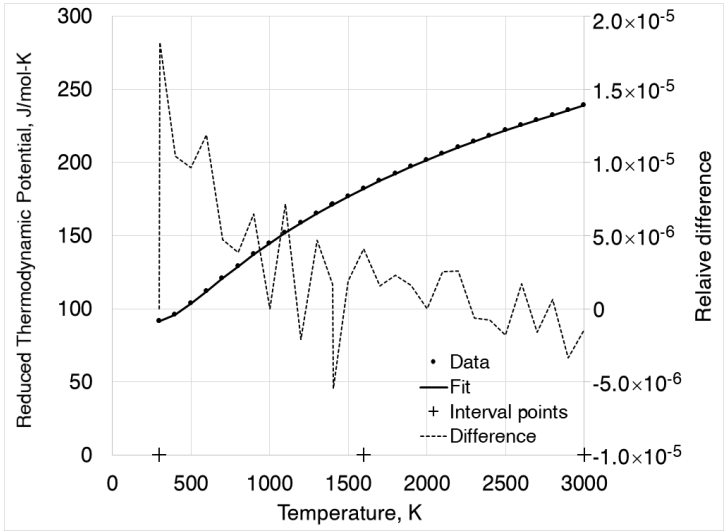
Sodium Oxide (Na_2O) – Na2O1(l)

MW = 61.9789 g/mol – $\Delta_f H^o$(298.15 K) = -372.843 kJ/mol – S^o(298.15 K) = 91.607 J/mol-K – p^o = 0.1 MPa



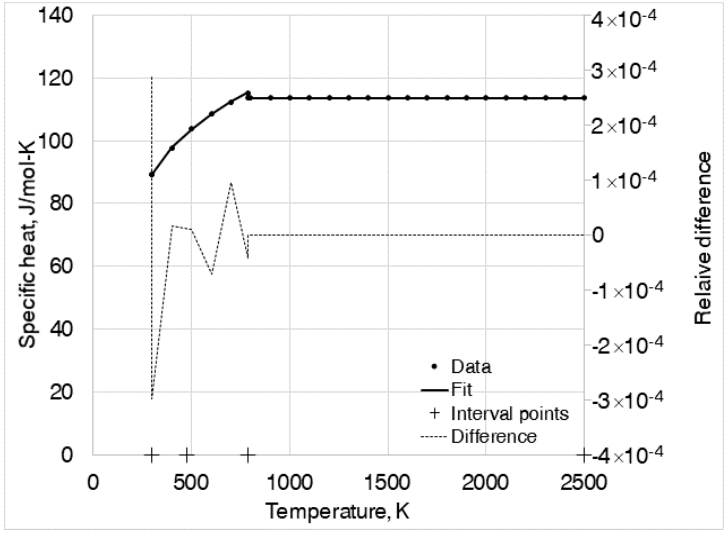


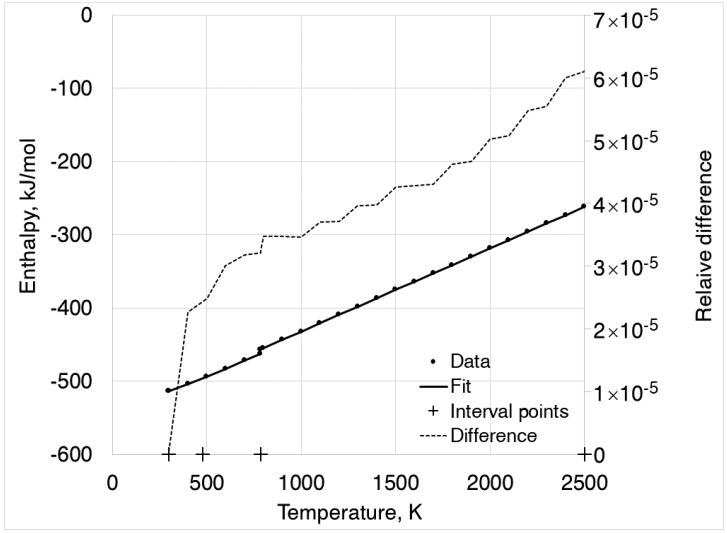


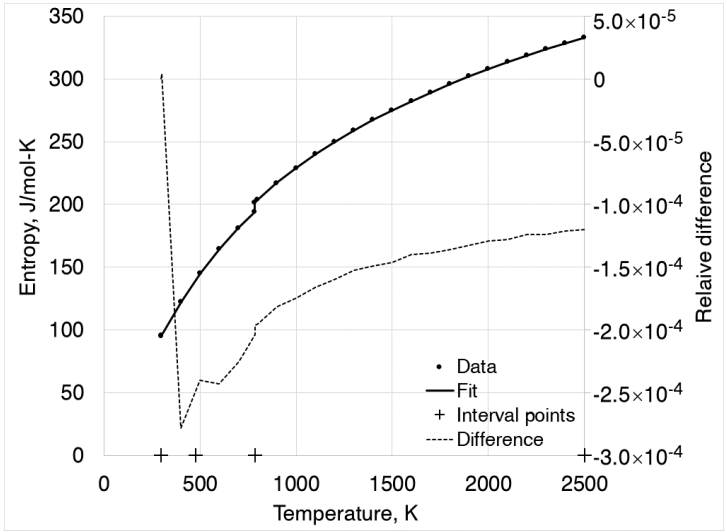


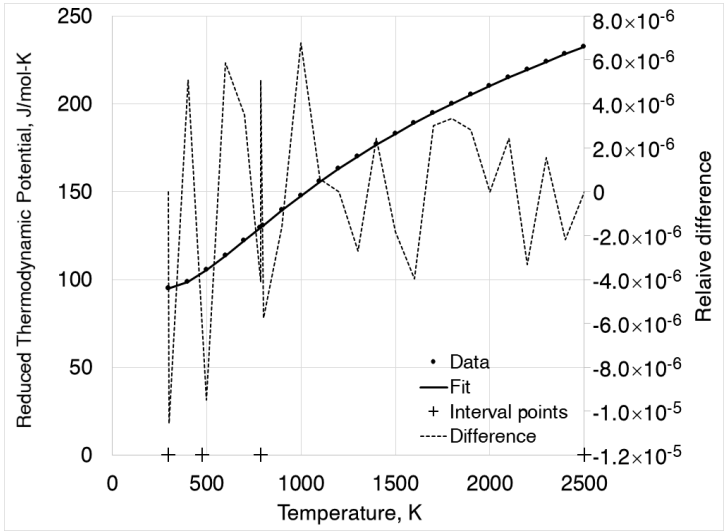
Sodium Peroxide (Na_2O_2) – Na2O2(cr)

MW = 77.9783 g/mol – $\Delta_f H^o$(298.15 K) = -513.209 kJ/mol – S^o(298.15 K) = 94.801 J/mol-K – p^o = 0.1 MPa



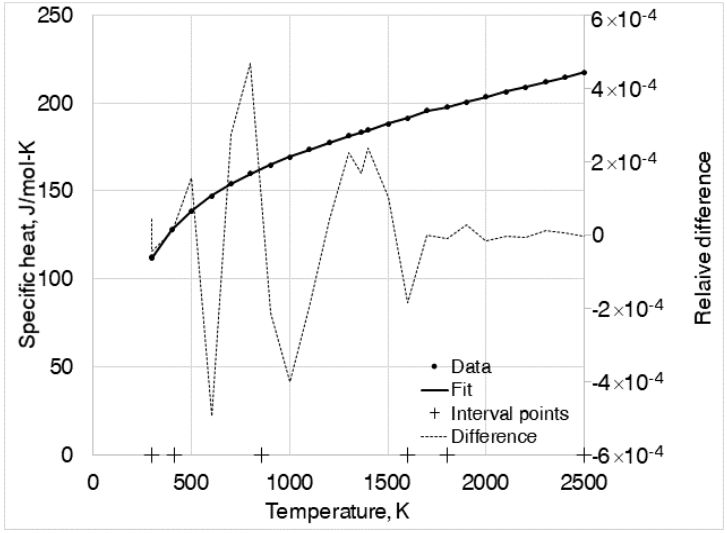


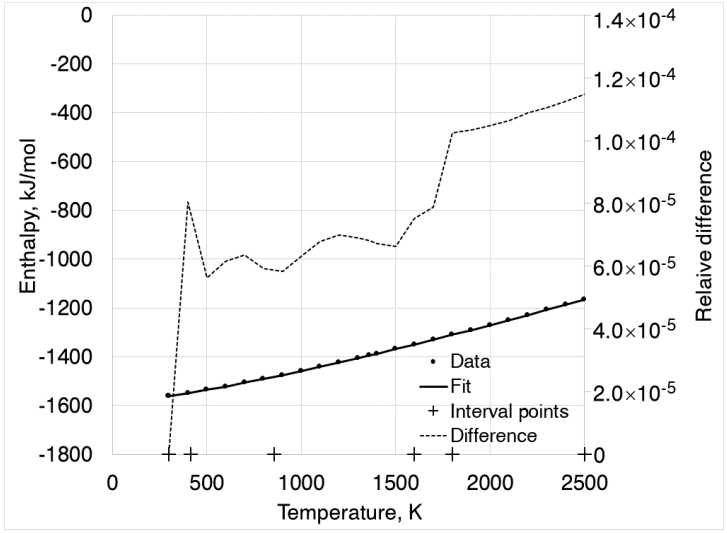


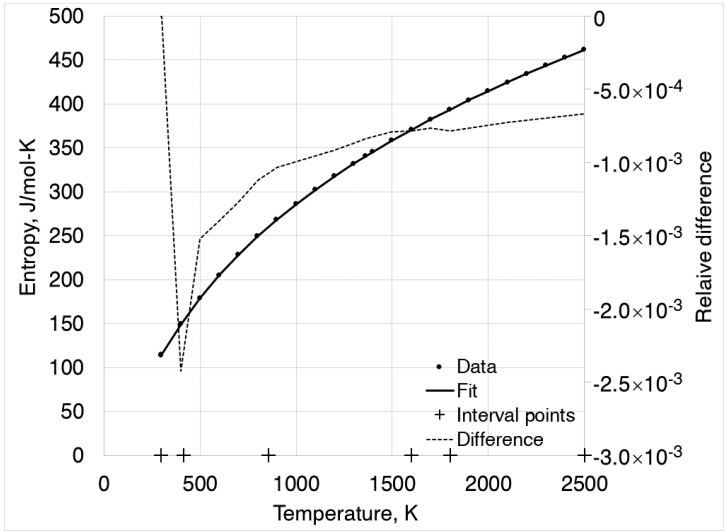


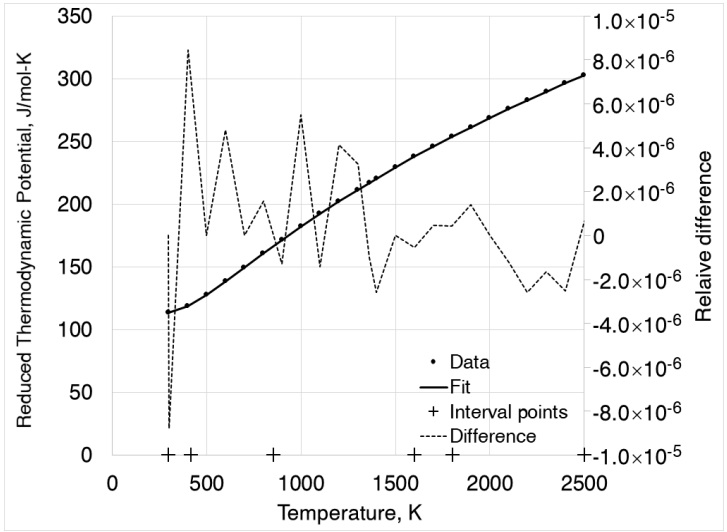
Sodium Silicate (Na_2SiO_3) – Na2O3Si1(cr)

MW = 122.0632 g/mol – $\Delta_f H^o$(298.15 K) = -1561.427 kJ/mol – S^o(298.15 K) = 113.847 J/mol-K – p^o = 0.1 MPa



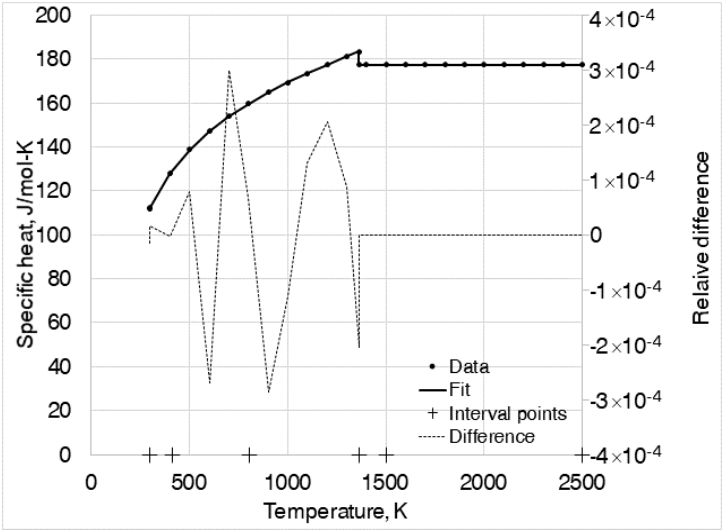


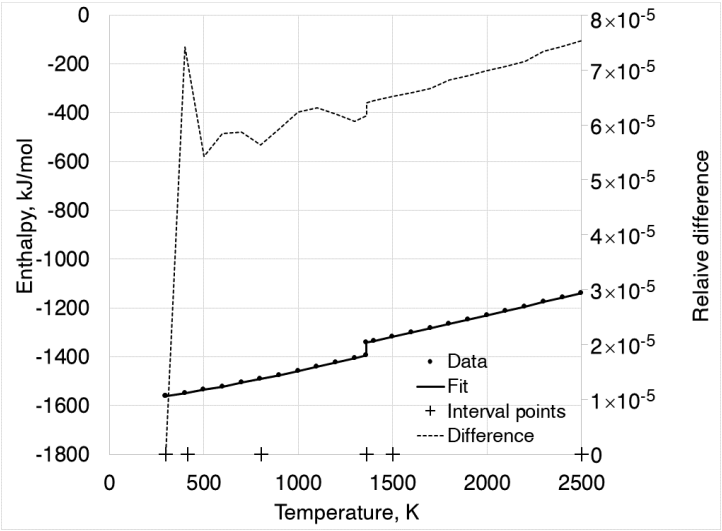


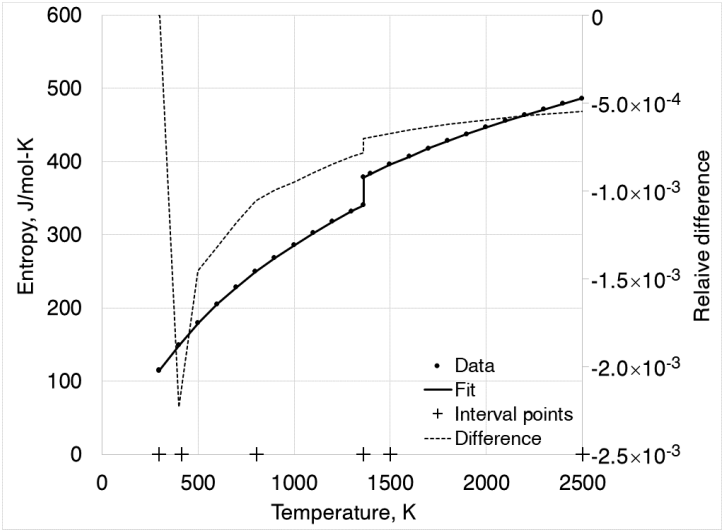


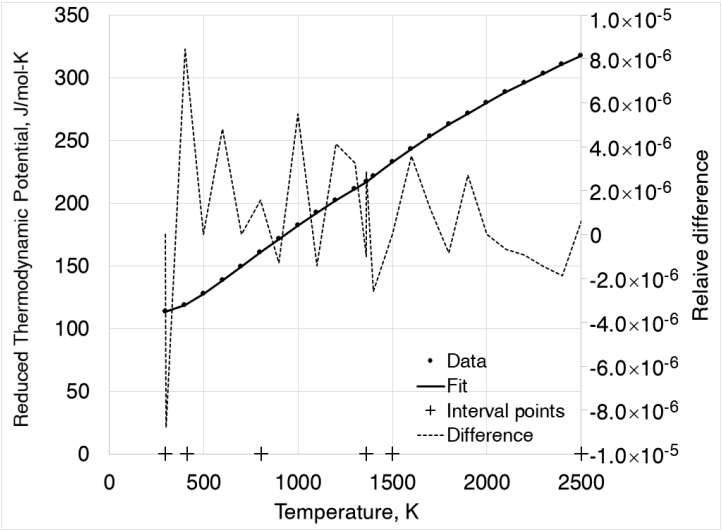
Sodium Silicate (Na_2SiO_3) – Na2O3Si1(cr,l)

MW = 122.0632 g/mol – $\Delta_f H^o$(298.15 K) = -1561.427 kJ/mol – S^o(298.15 K) = 113.847 J/mol-K – p^o = 0.1 MPa



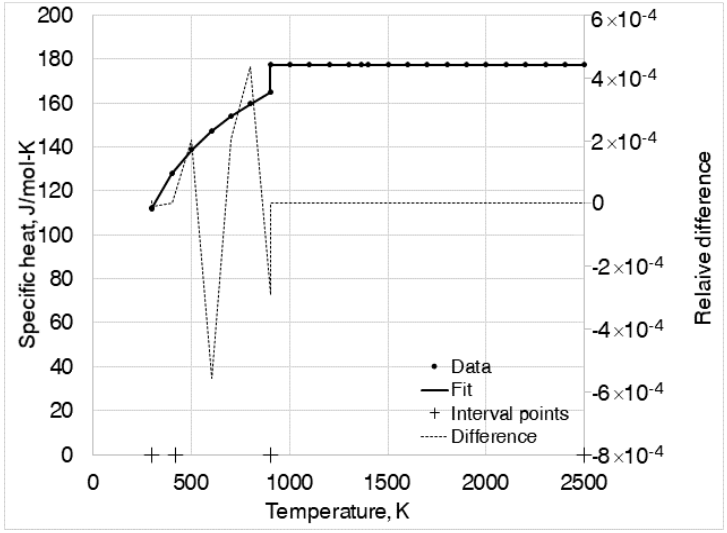


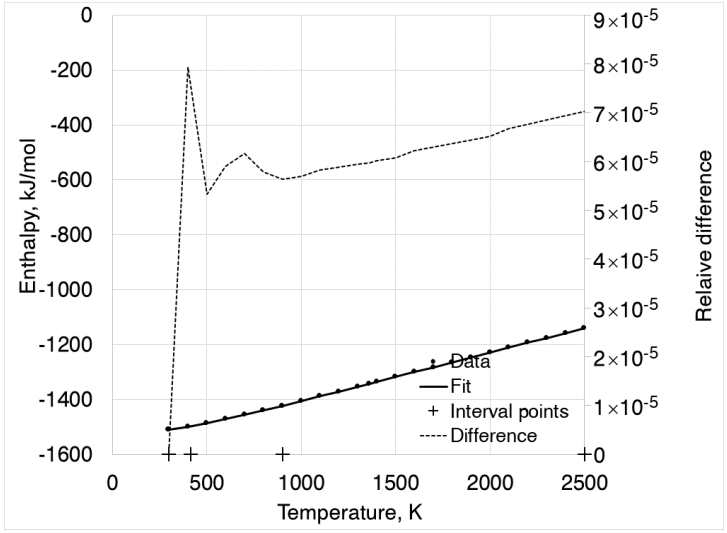


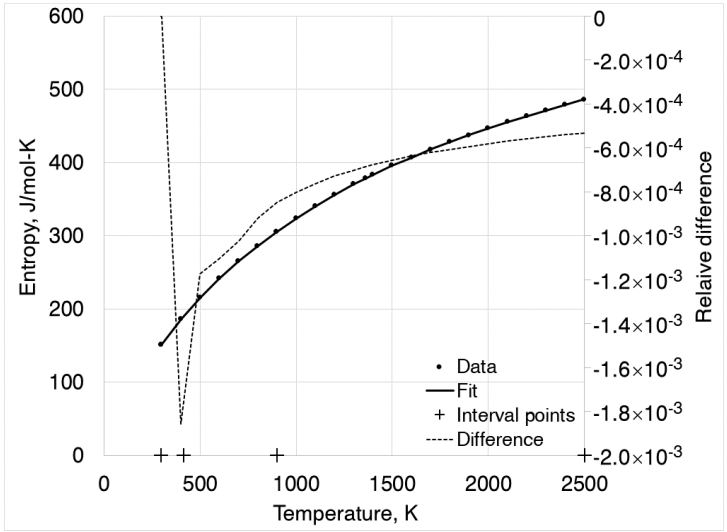


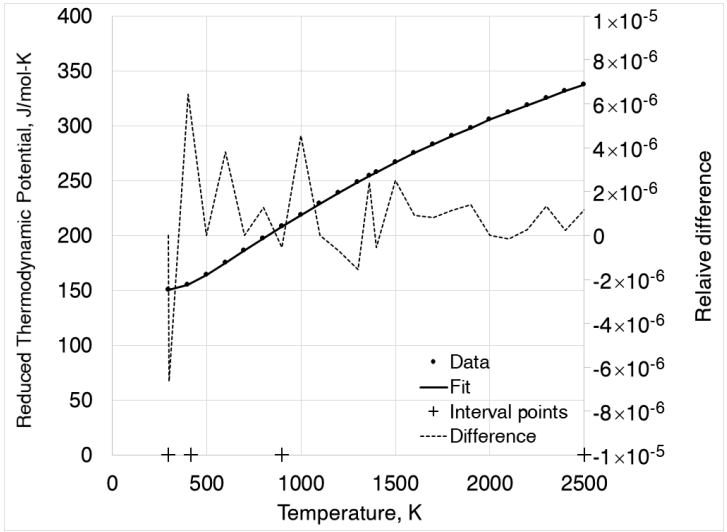
Sodium Silicate (Na_2SiO_3) – Na2O3Si1(l)

MW = 122.0632 g/mol – $\Delta_f H^o$(298.15 K) = -1510.875 kJ/mol – S^o(298.15 K) = 150.496 J/mol-K – p^o = 0.1 MPa



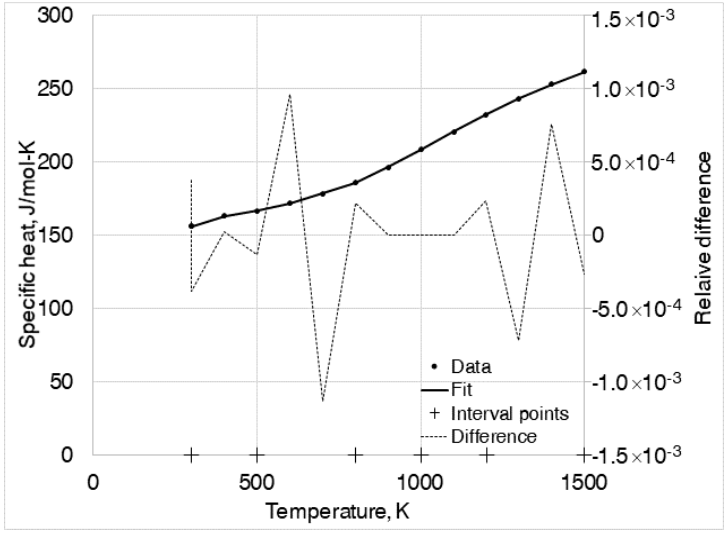


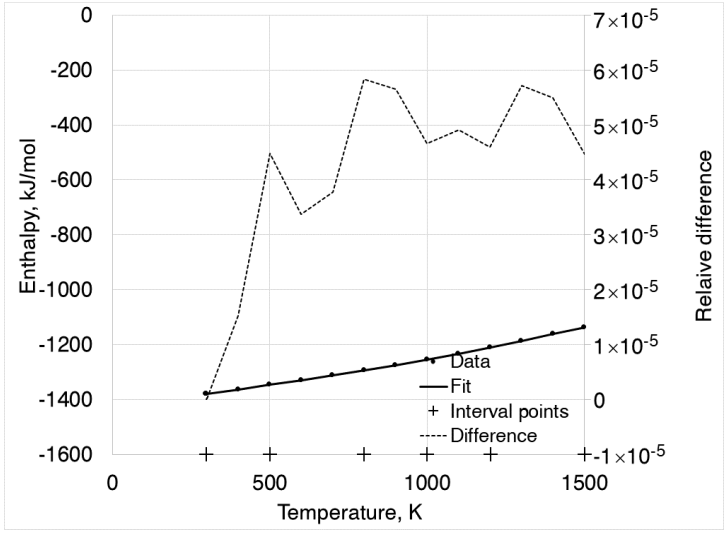


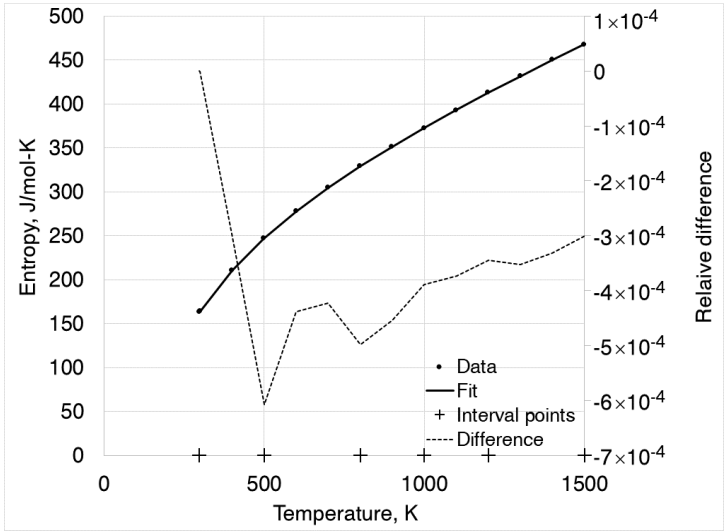


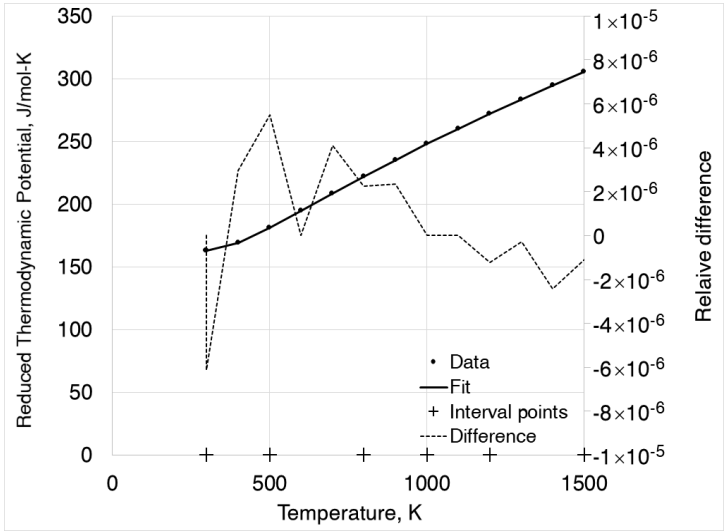
Sodium Sulfate, Delta (Na_2SO_4) – Na2O4S1(cr)_01

MW = 142.04 g/mol – $\Delta_f H^o$(298.15 K) = -1379.294 kJ/mol – S^o(298.15 K) = 163.212 J/mol-K – p^o = 0.1 MPa



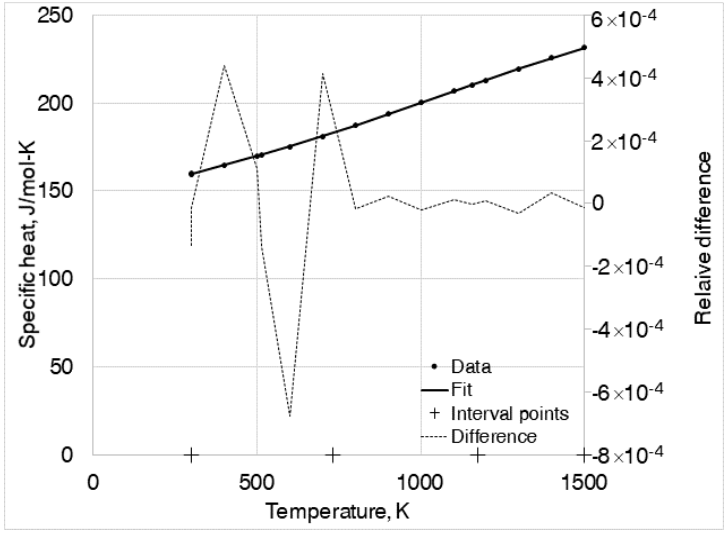


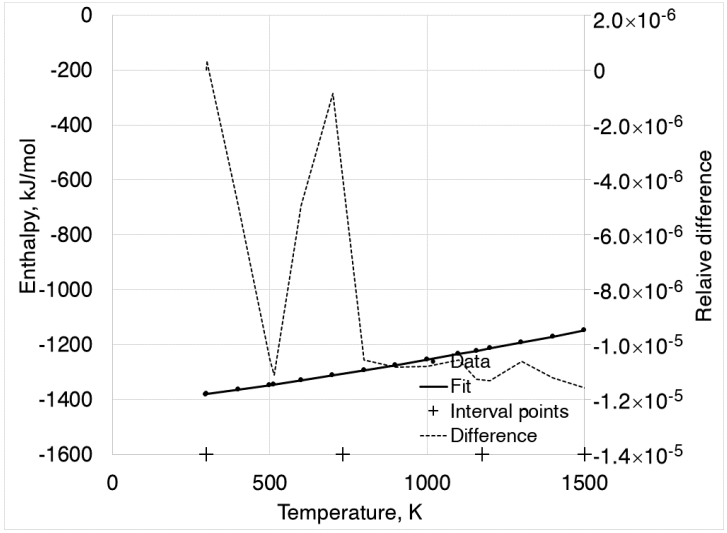


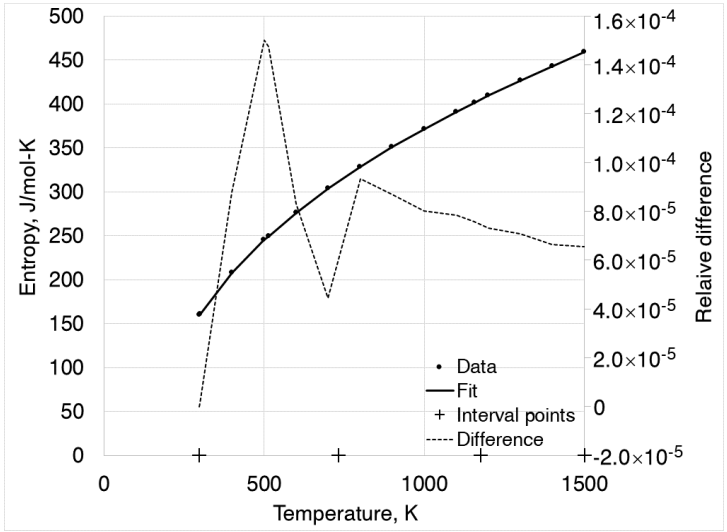


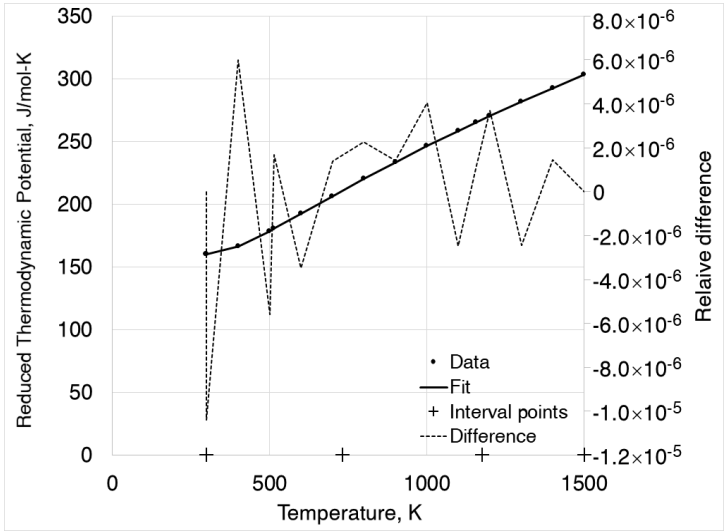
Sodium Sulfate, I (Na_2SO_4) – Na2O4S1(cr)_02

MW = 142.04 g/mol – $\Delta_f H^o$(298.15 K) = -1380.899 kJ/mol – S^o(298.15 K) = 160.132 J/mol-K – p^o = 0.1 MPa



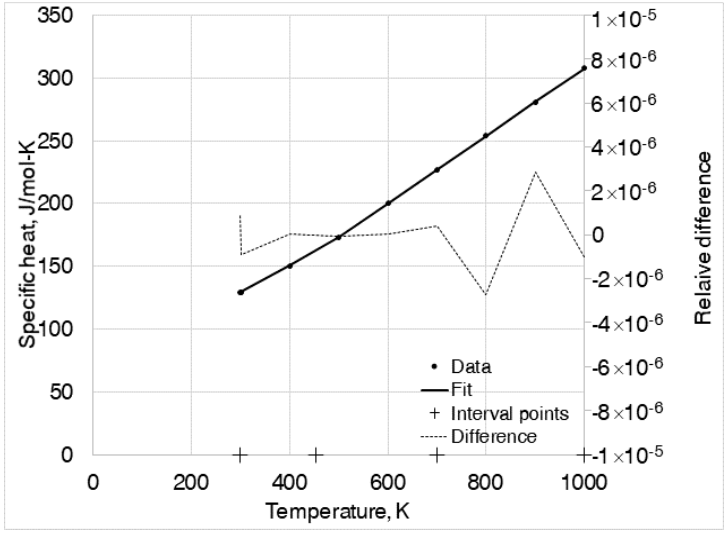


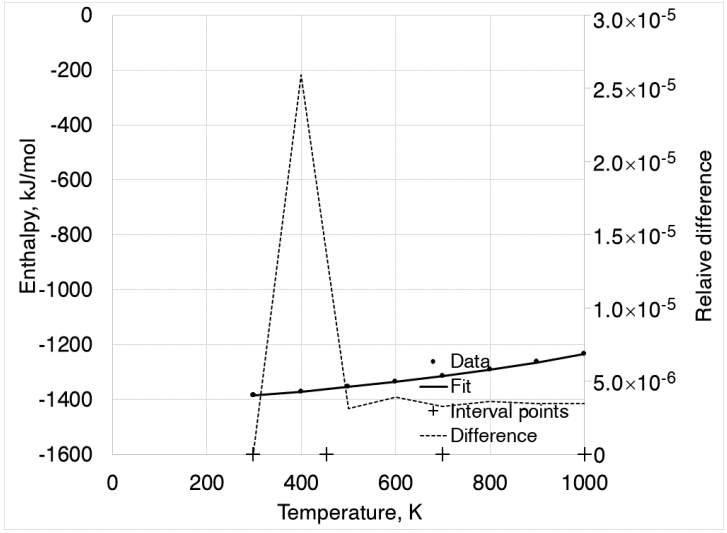


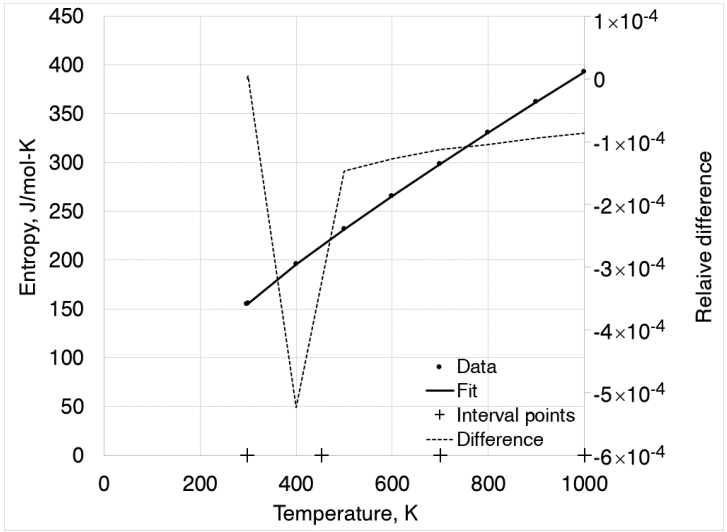


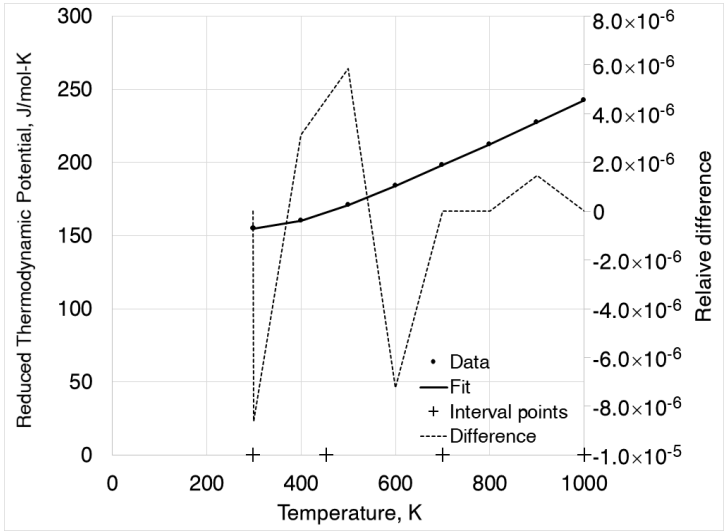
Sodium Sulfate, III (Na_2SO_4) – Na2O4S1(cr)_03

MW = 142.04 g/mol – $\Delta_f H^o$(298.15 K) = -1384.854 kJ/mol – S^o(298.15 K) = 154.908 J/mol-K – p^o = 0.1 MPa



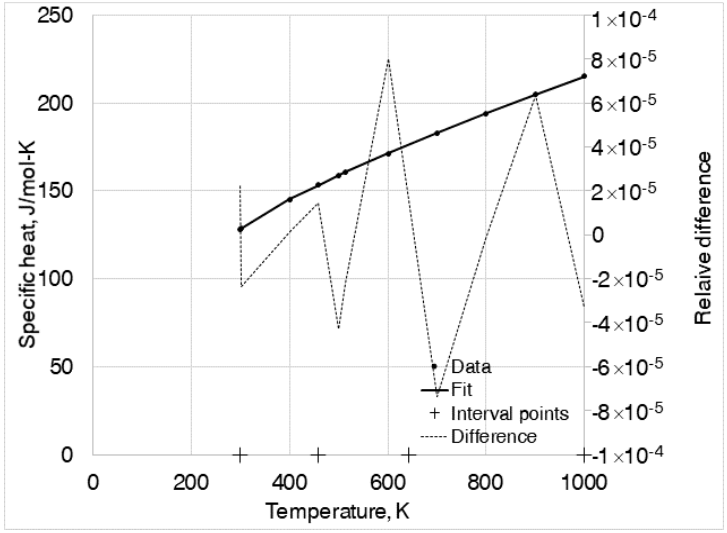


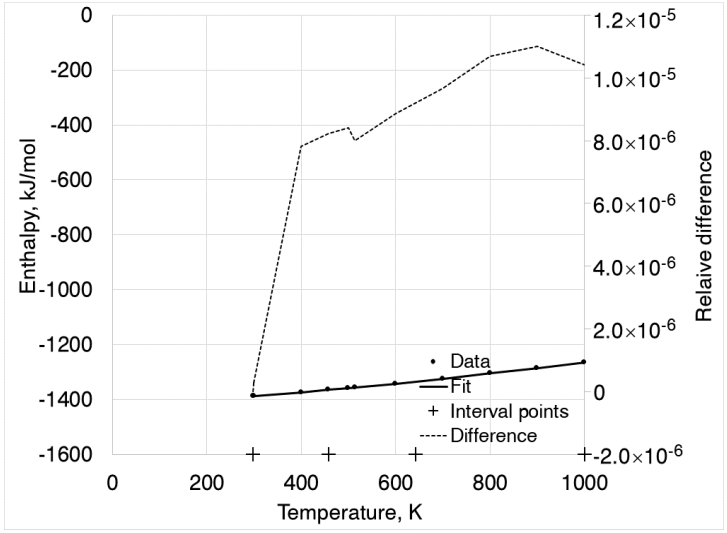


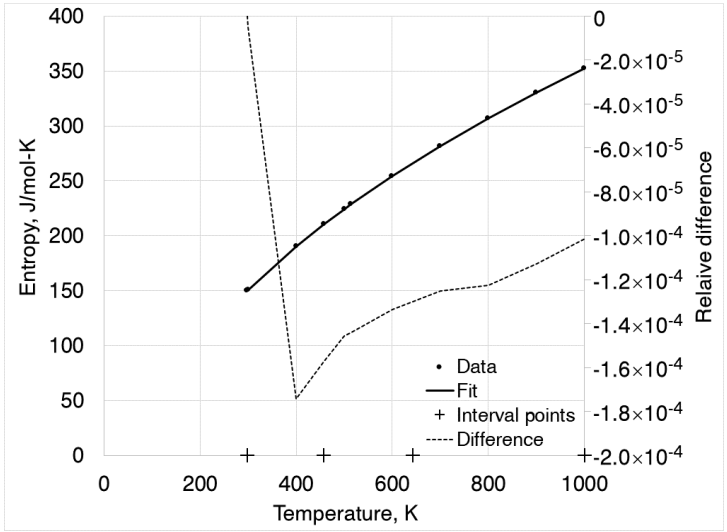


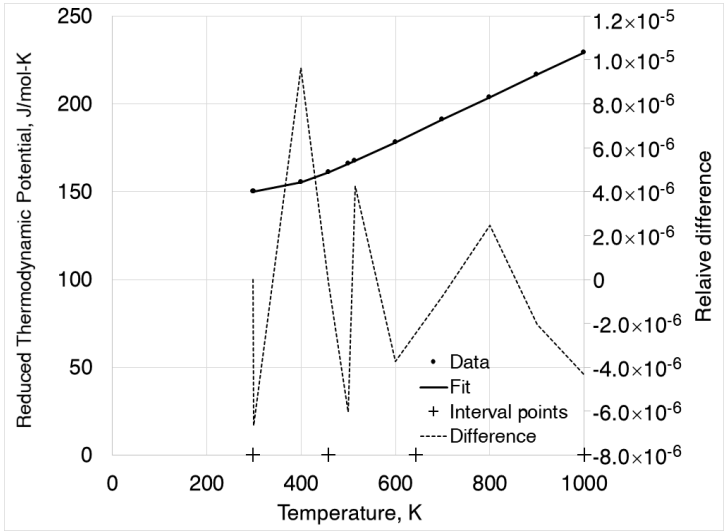
Sodium Sulfate, IV (Na_2SO_4) – Na2O4S1(cr)_04

MW = 142.04 g/mol – $\Delta_f H^o$(298.15 K) = -1387.559 kJ/mol – S^o(298.15 K) = 150.157 J/mol-K – p^o = 0.1 MPa



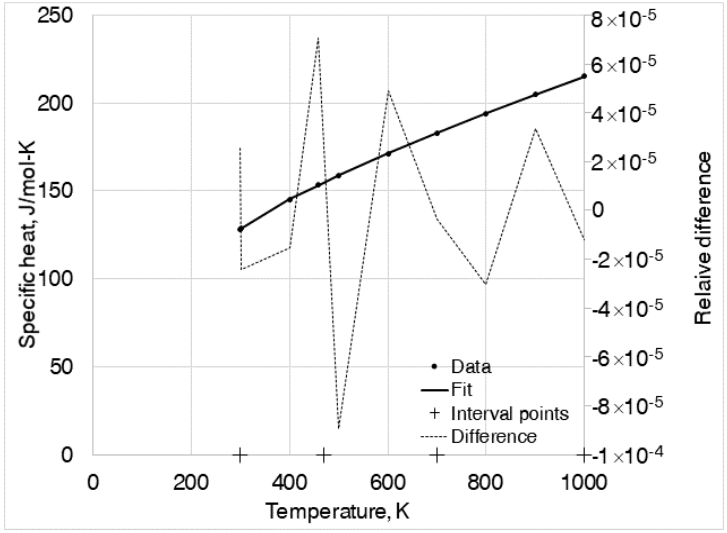


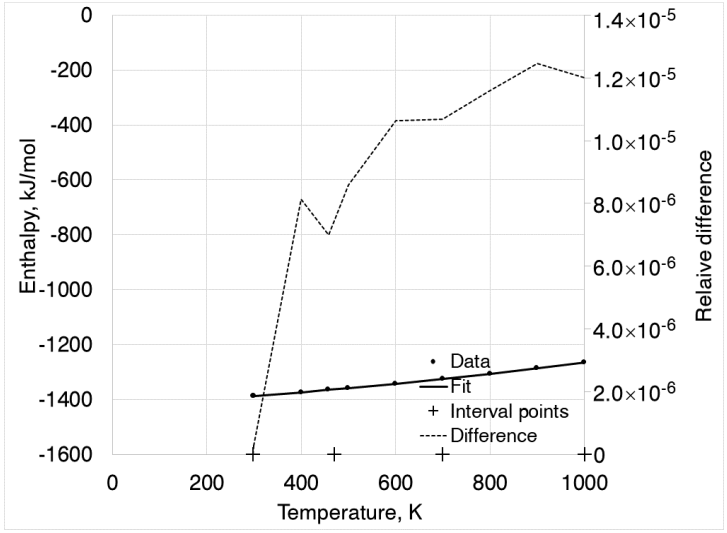


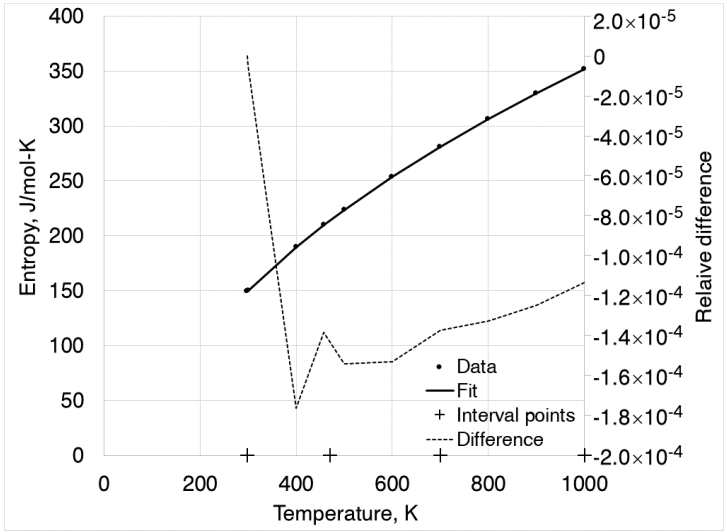


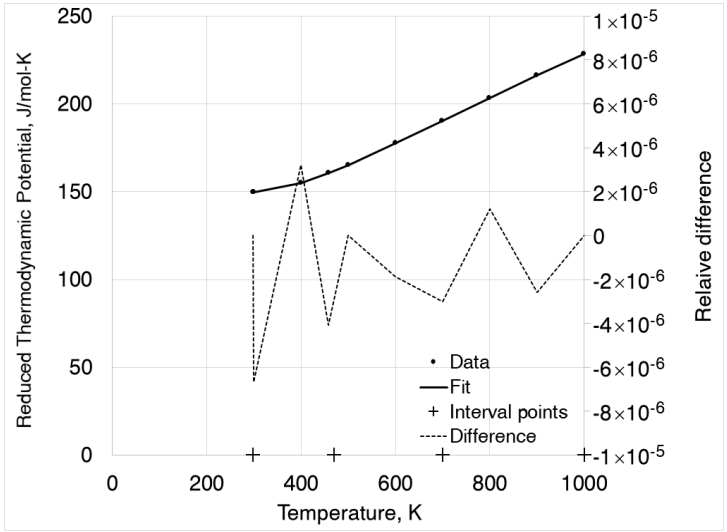
Sodium Sulfate, V (Na_2SO_4) – Na2O4S1(cr)_05

MW = 142.04 g/mol – $\Delta_f H^o$(298.15 K) = -1387.816 kJ/mol – S^o(298.15 K) = 149.595 J/mol-K – p^o = 0.1 MPa



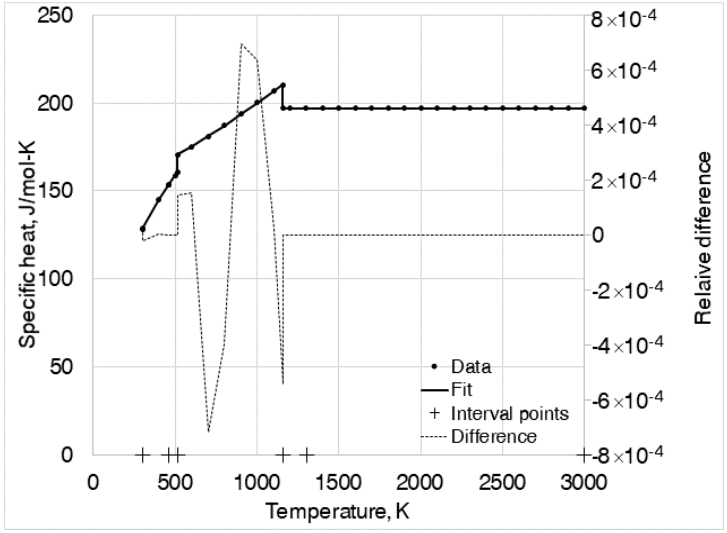


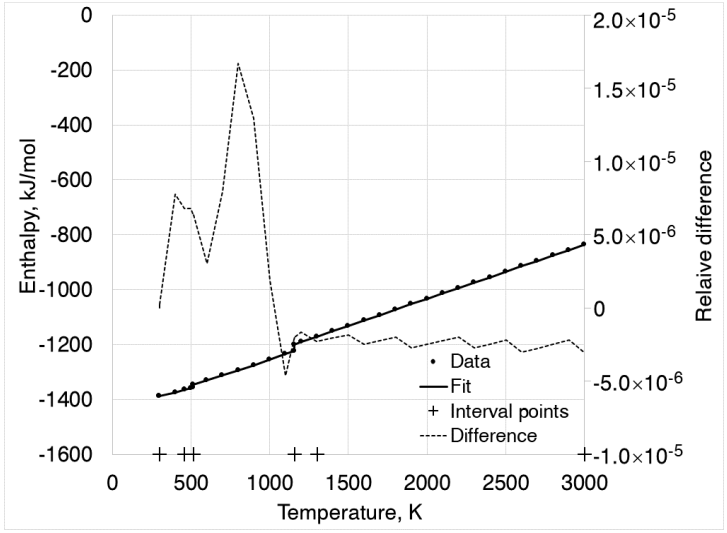


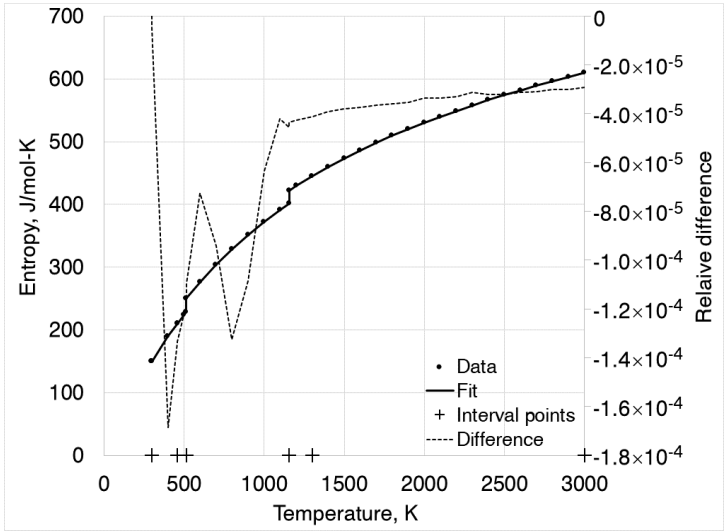


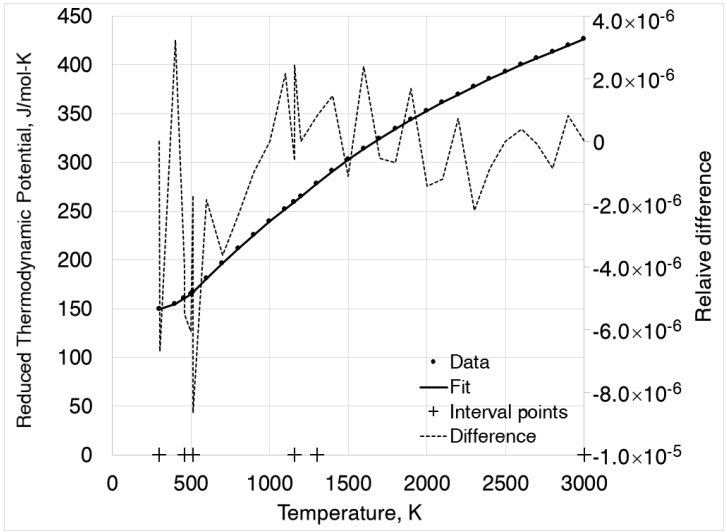
Sodium Sulfate (Na_2SO_4) – Na2O4S1(cr,l)

MW = 142.04 g/mol – $\Delta_f H^o$(298.15 K) = -1387.816 kJ/mol – S^o(298.15 K) = 149.595 J/mol-K – p^o = 0.1 MPa



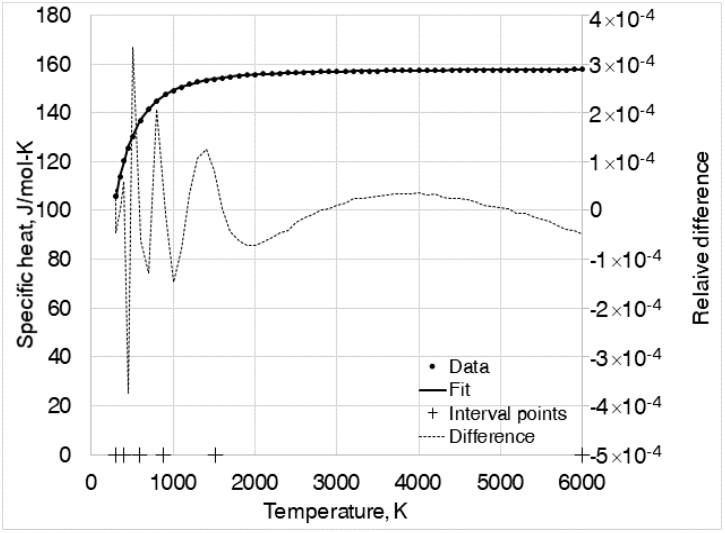


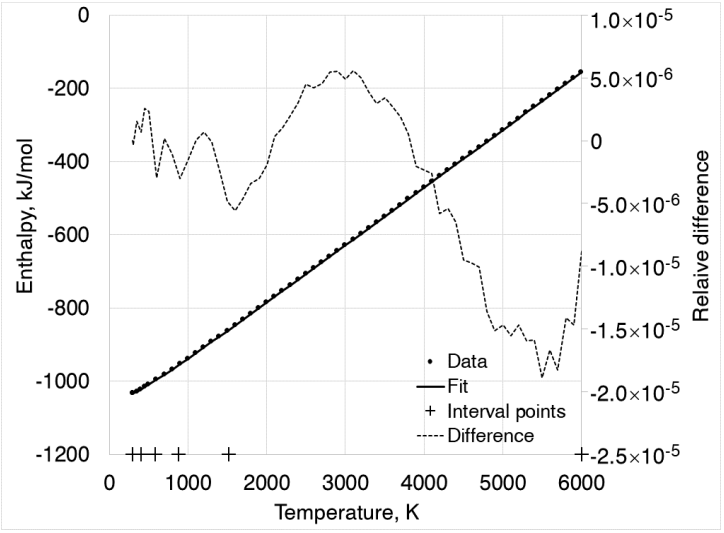


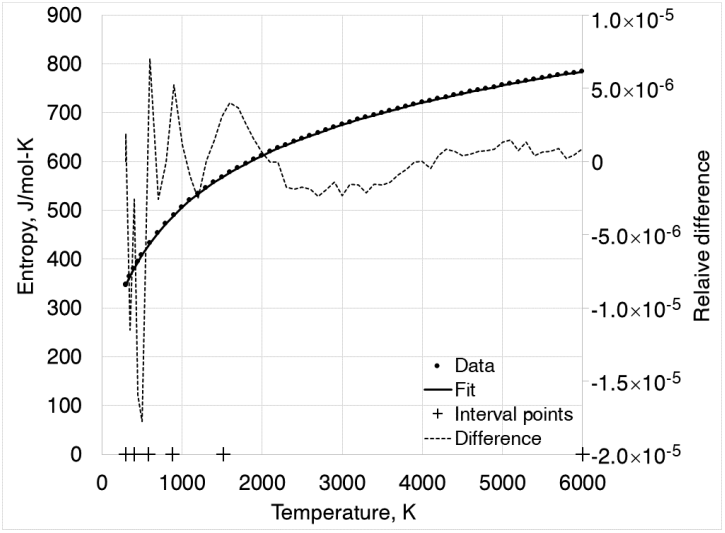


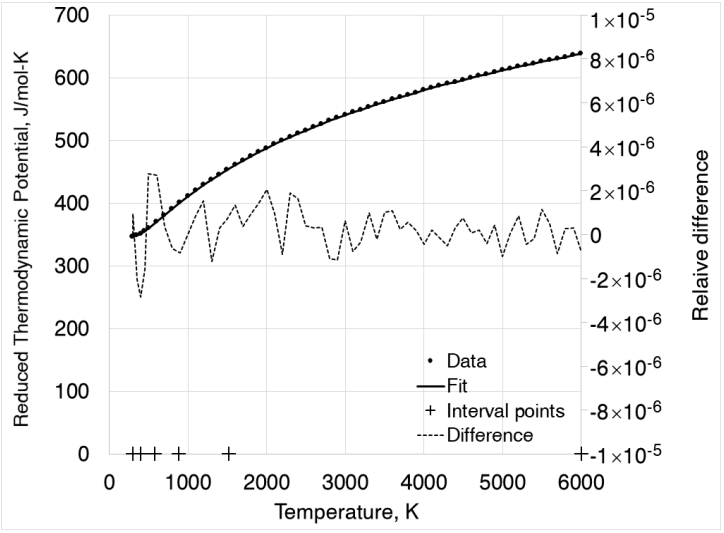
Sodium Sulfate (Na_2SO_4) – Na2O4S1(g)

MW = 142.04 g/mol – $\Delta_f H^o$(298.15 K) = -1033.615 kJ/mol – S^o(298.15 K) = 346.854 J/mol-K – p^o = 0.1 MPa



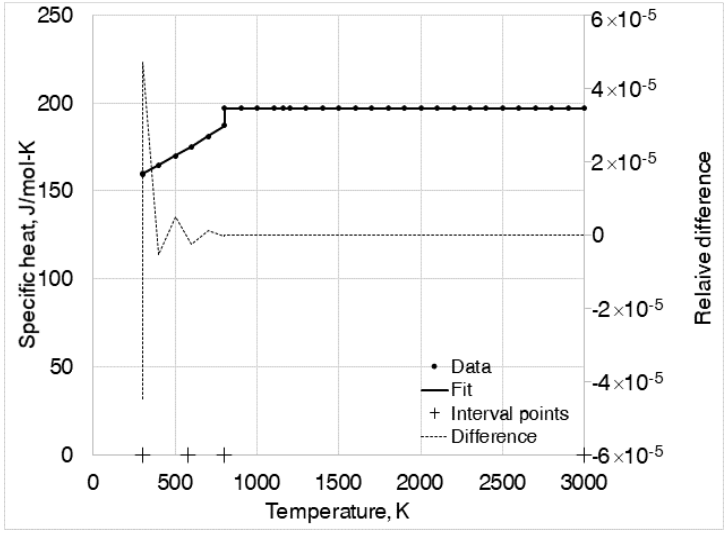


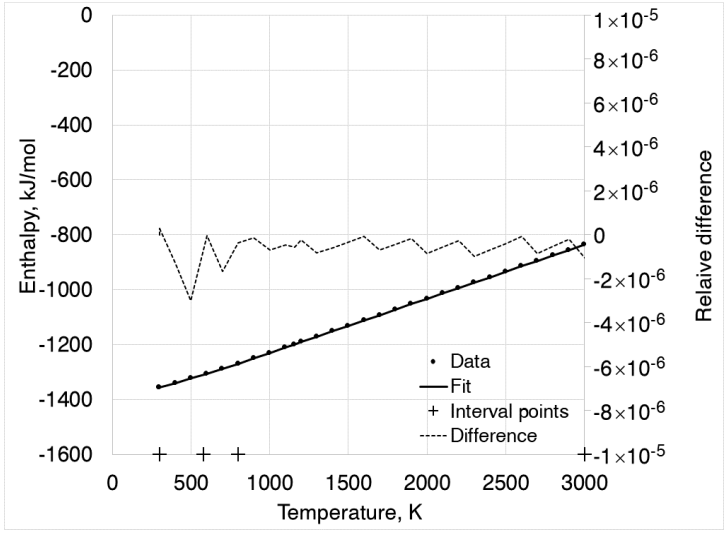


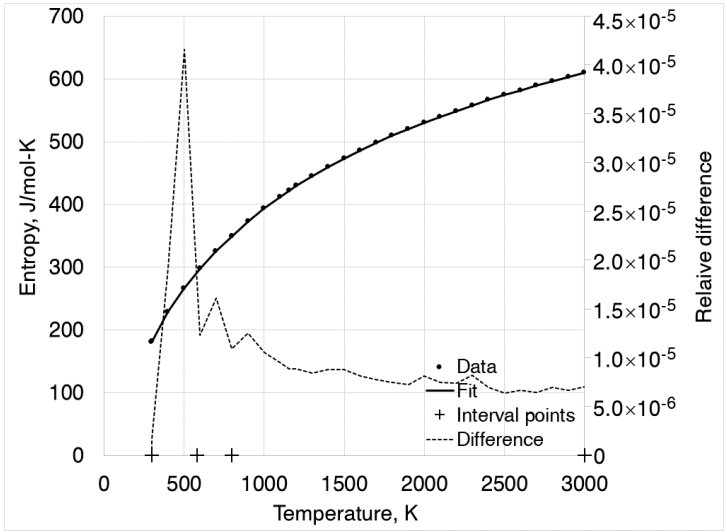


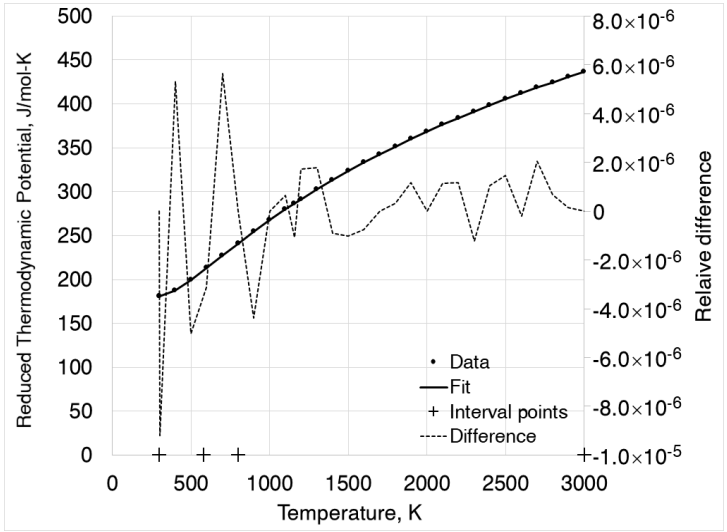
Sodium Sulfate (Na_2SO_4) – Na2O4S1(l)

MW = 142.04 g/mol – $\Delta_f H^o$(298.15 K) = -1356.385 kJ/mol – S^o(298.15 K) = 181.172 J/mol-K – p^o = 0.1 MPa



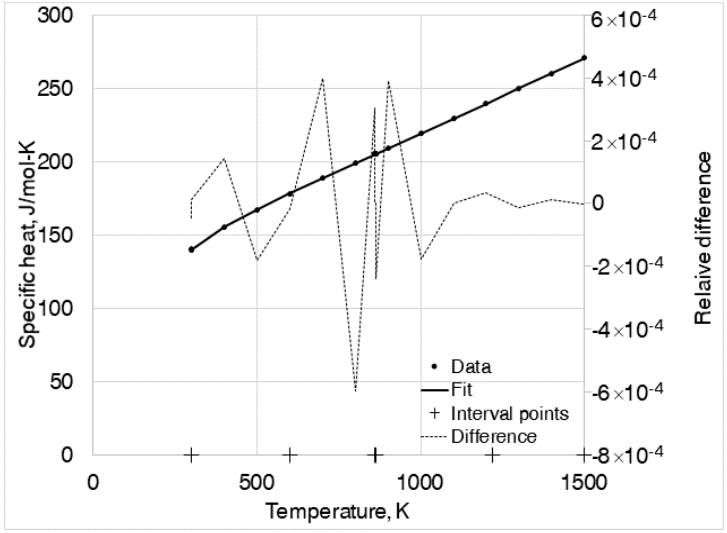


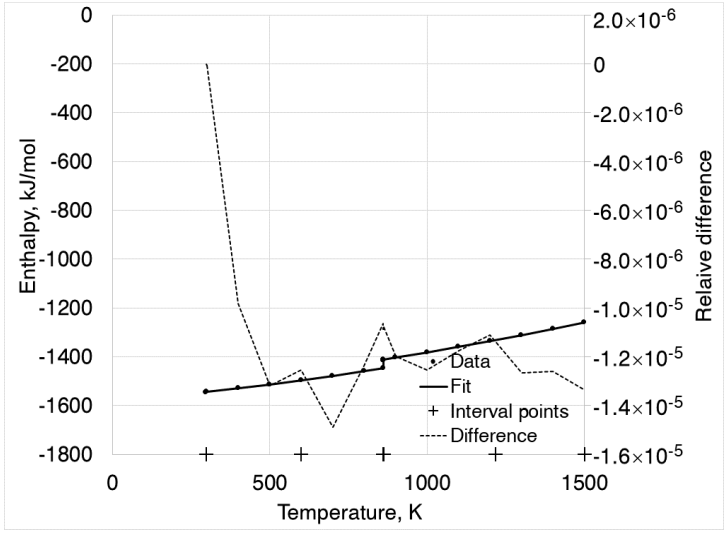


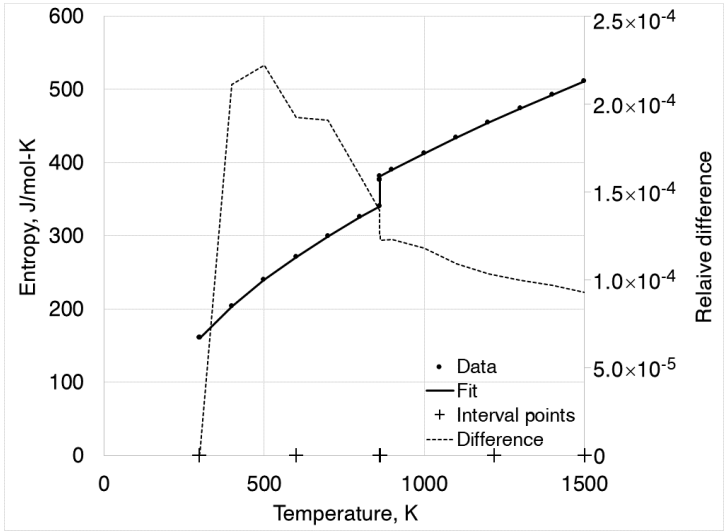


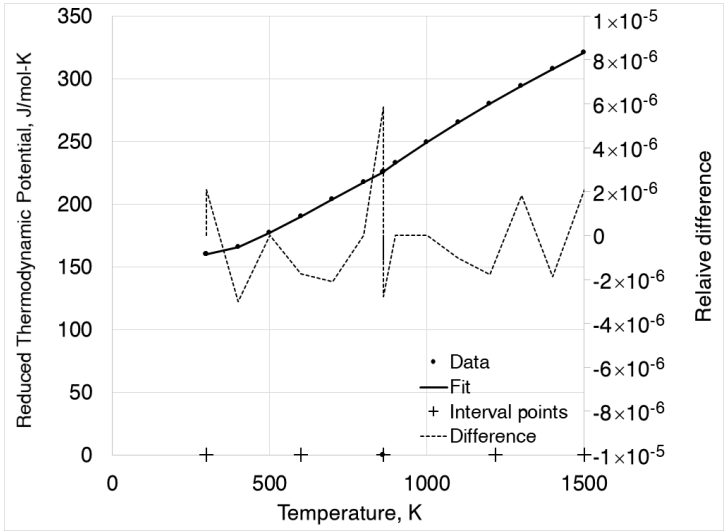
Sodium Tungsten Oxide (Na_2WO_4) – Na2O4W1(cr)

MW = 293.83 g/mol – $\Delta_f H^o$(298.15 K) = -1544.733 kJ/mol – S^o(298.15 K) = 160.331 J/mol-K – p^o = 0.1 MPa



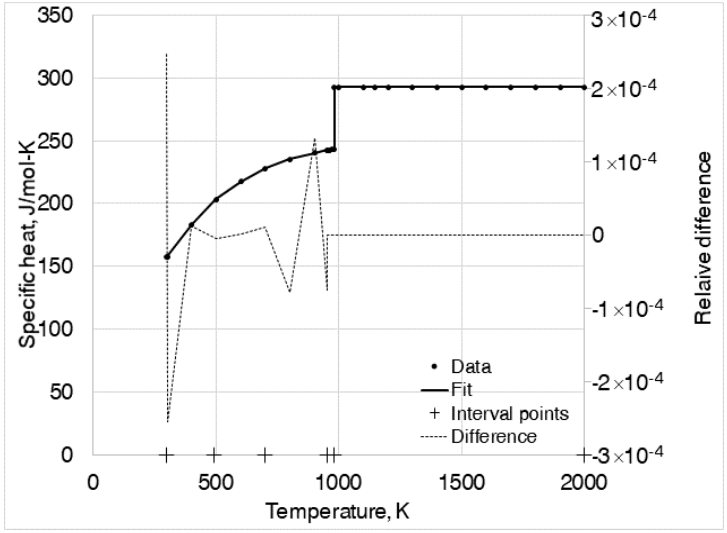


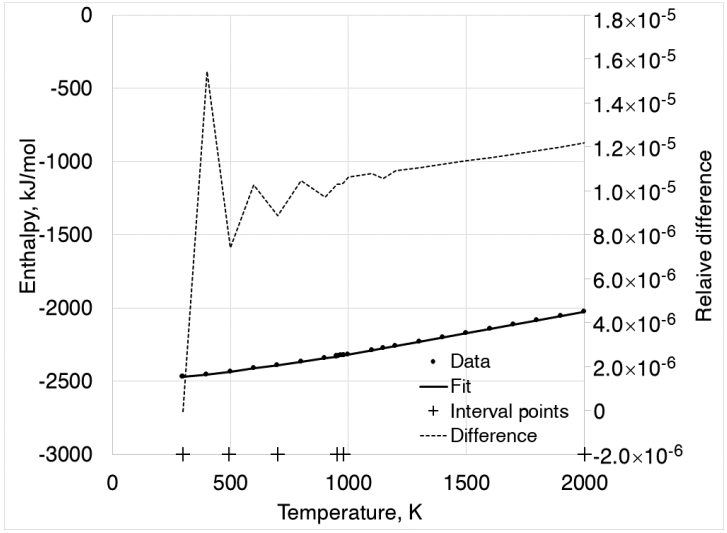


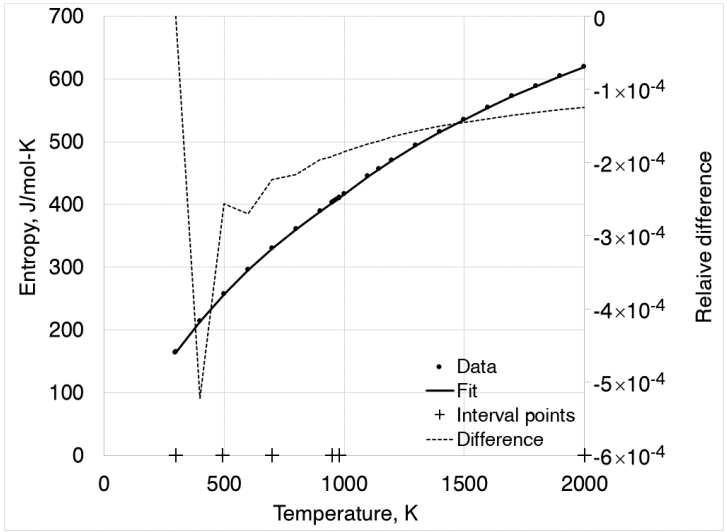


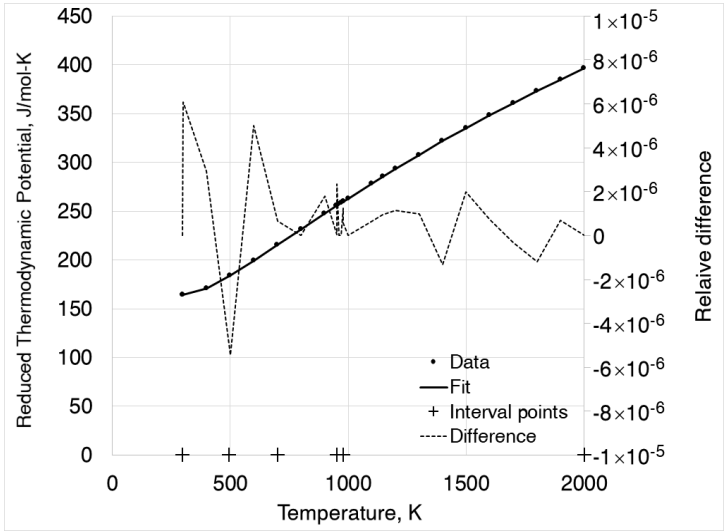
Sodium Silicate ($Na_2Si_2O_5$) – Na2O5Si2(cr)

MW = 182.1475 g/mol – $\Delta_f H^o$(298.15 K) = -2470.066 kJ/mol – S^o(298.15 K) = 164.055 J/mol-K – p^o = 0.1 MPa



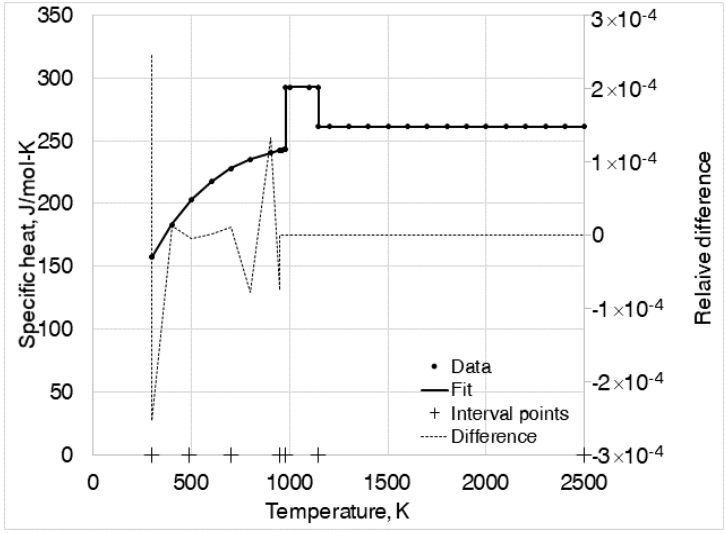


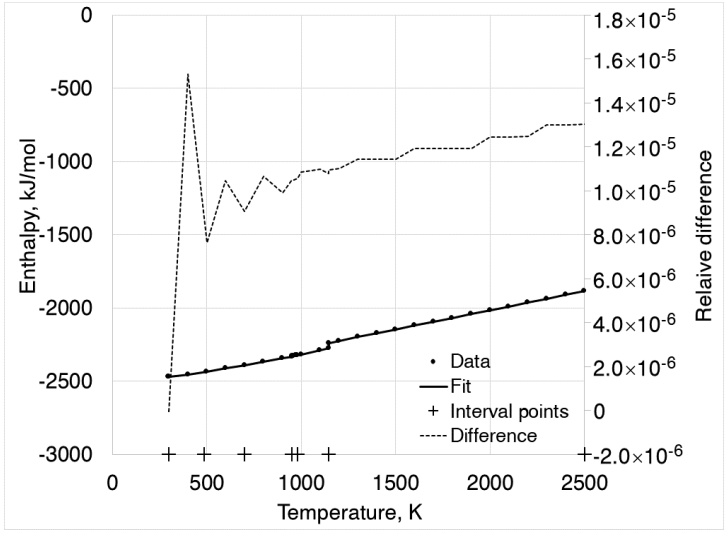


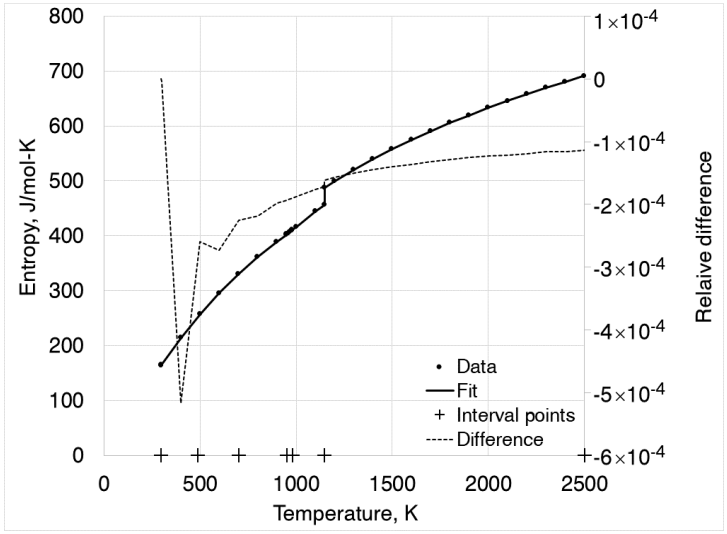


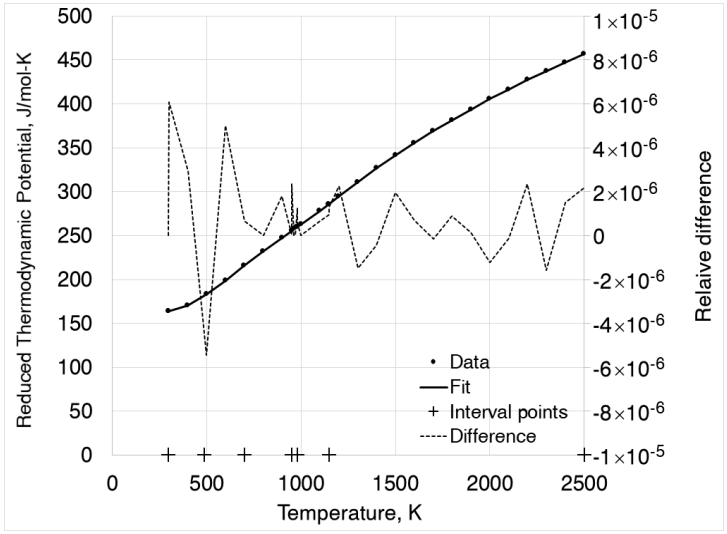
Sodium Silicate ($Na_2Si_2O_5$) – Na2O5Si2(cr,l)

MW = 182.1475 g/mol – $\Delta_f H^o$(298.15 K) = -2470.066 kJ/mol – S^o(298.15 K) = 164.055 J/mol-K – p^o = 0.1 MPa



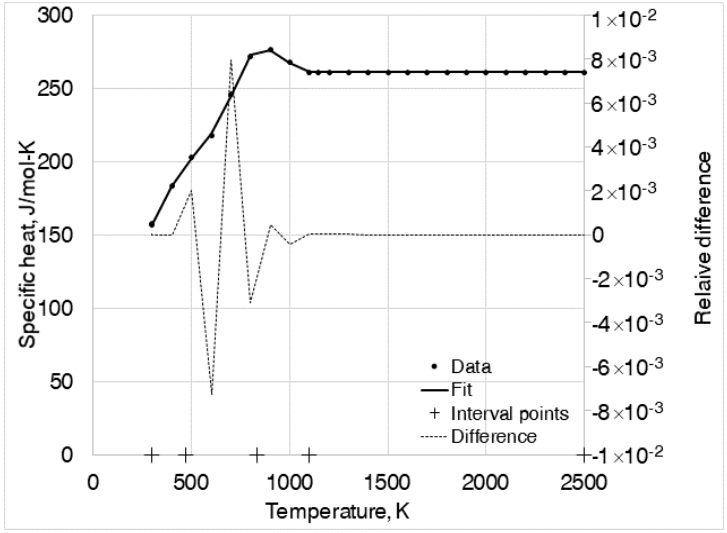


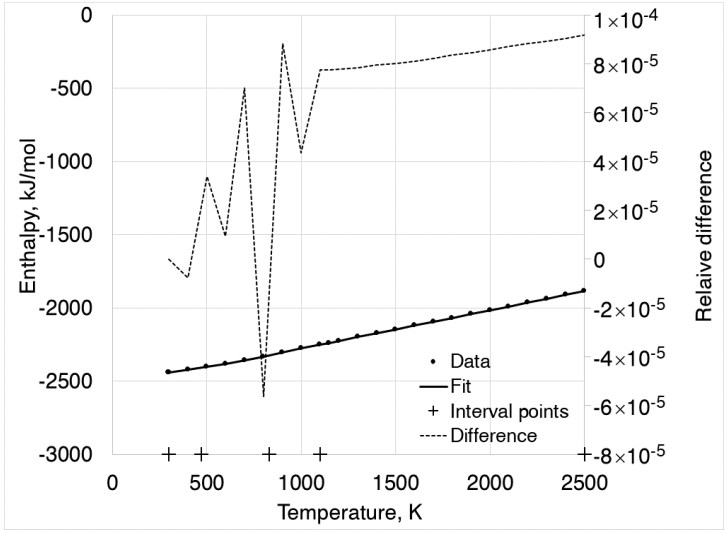


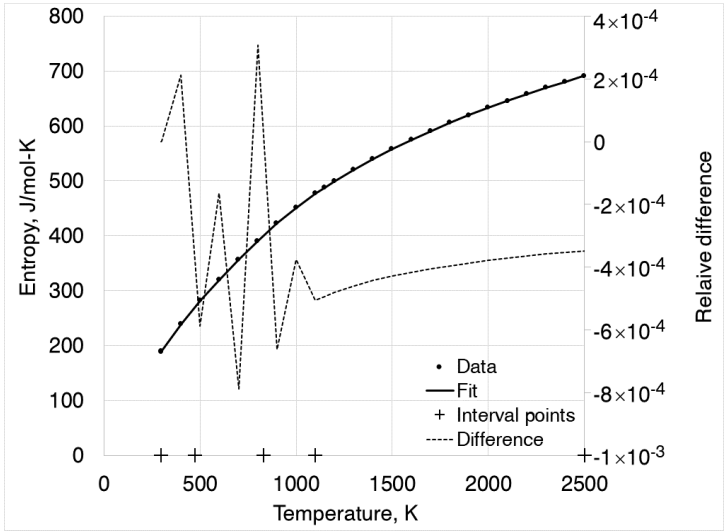


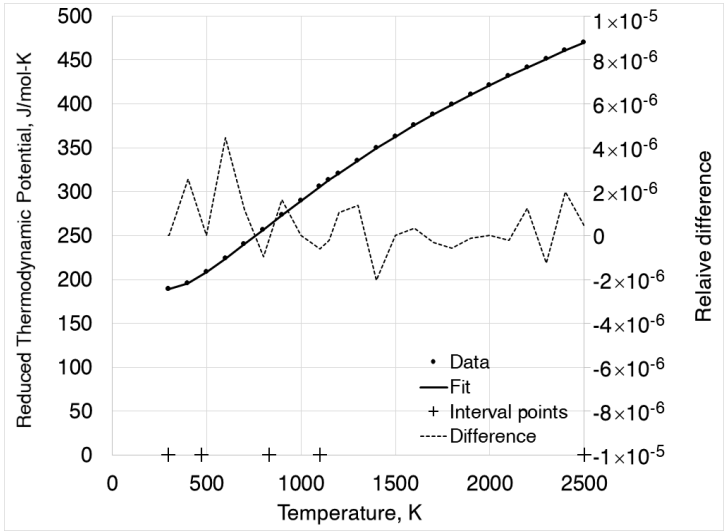
Sodium Silicate ($Na_2Si_2O_5$) – Na2O5Si2(l)

MW = 182.1475 g/mol – $\Delta_f H^o$(298.15 K) = -2438.485 kJ/mol – S^o(298.15 K) = 188.705 J/mol-K – p^o = 0.1 MPa



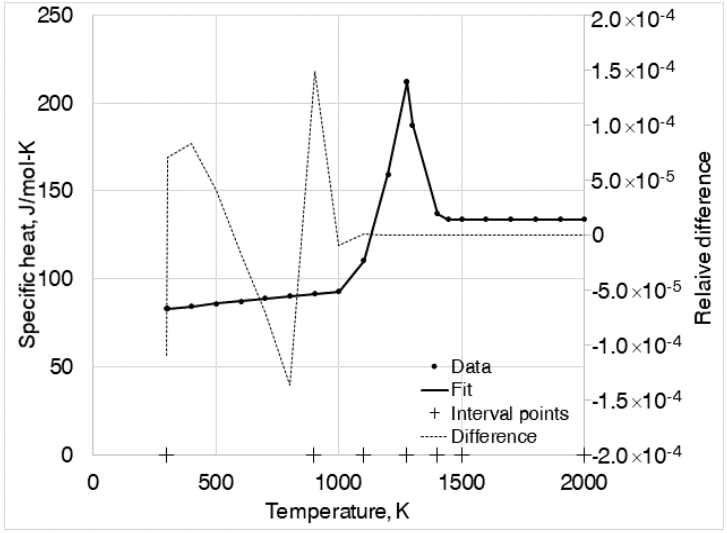


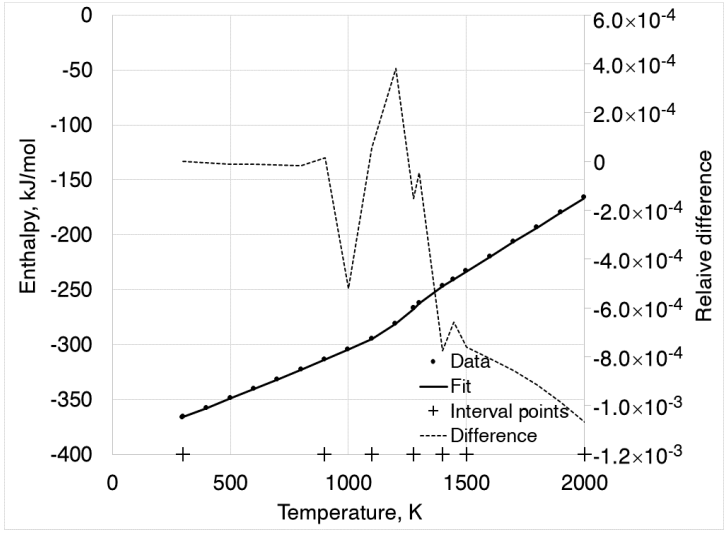


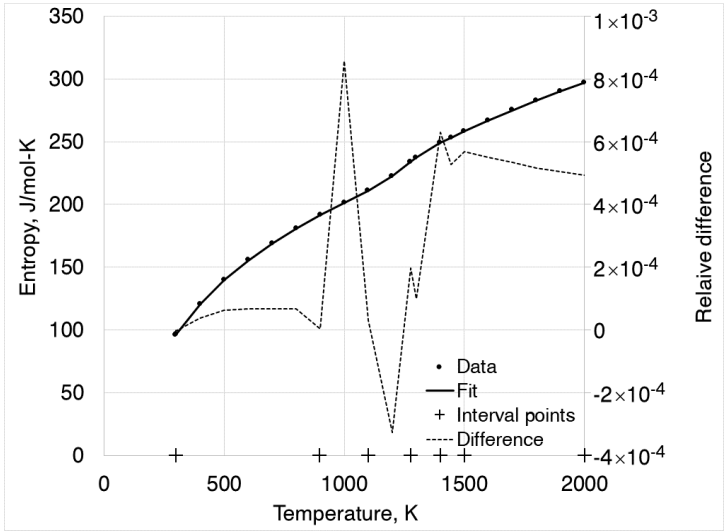


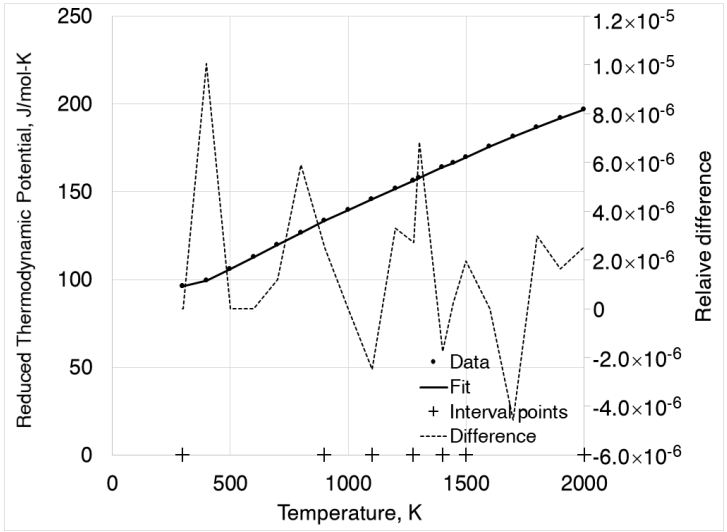
Sodium Sulfide (Na_2S) – Na2S1(cr)

MW = 78.04 g/mol – $\Delta_f H^o$(298.15 K) = -366.1 kJ/mol – S^o(298.15 K) = 96.232 J/mol-K – p^o = 0.1 MPa



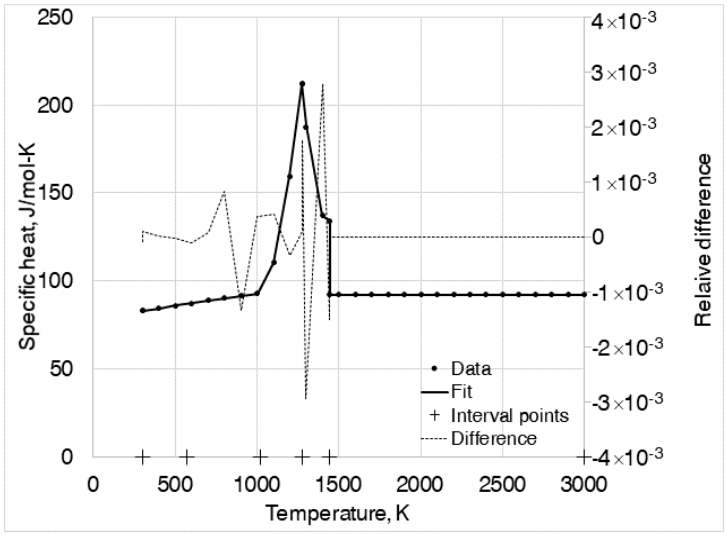


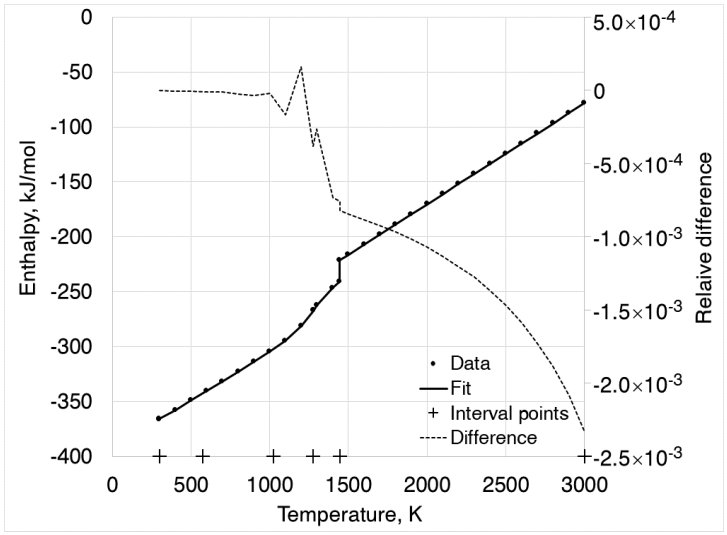


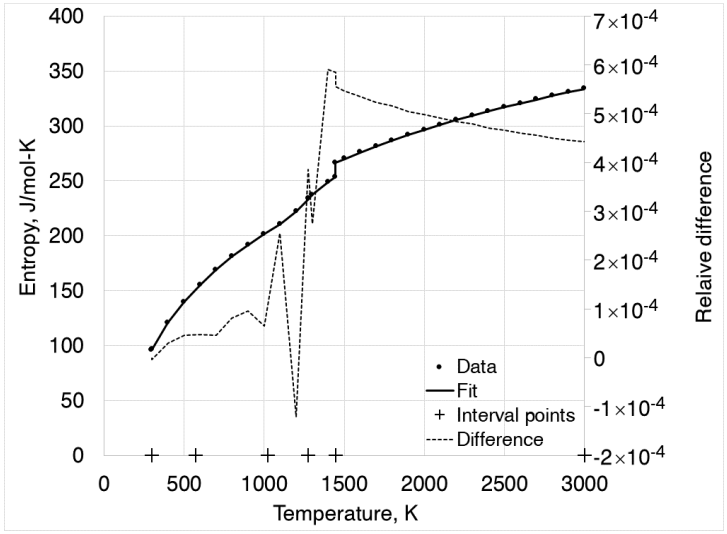


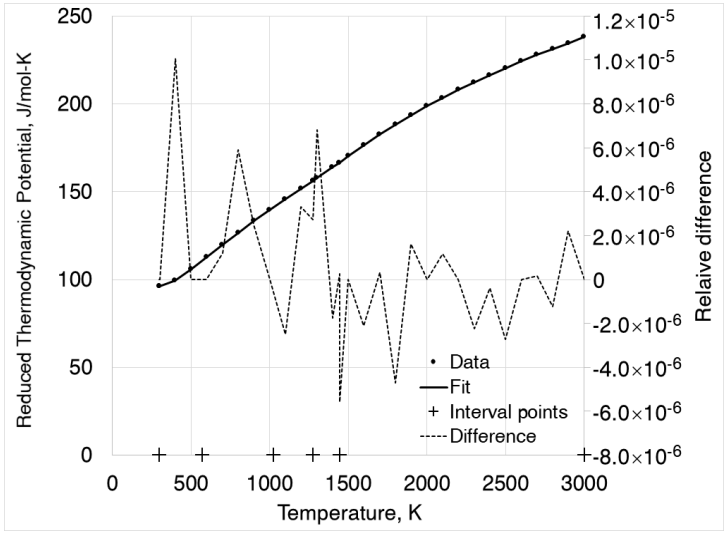
Sodium Sulfide (Na_2S) – Na2S1(cr,l)

MW = 78.04 g/mol – $\Delta_f H^o$(298.15 K) = -366.1 kJ/mol – S^o(298.15 K) = 96.232 J/mol-K – p^o = 0.1 MPa



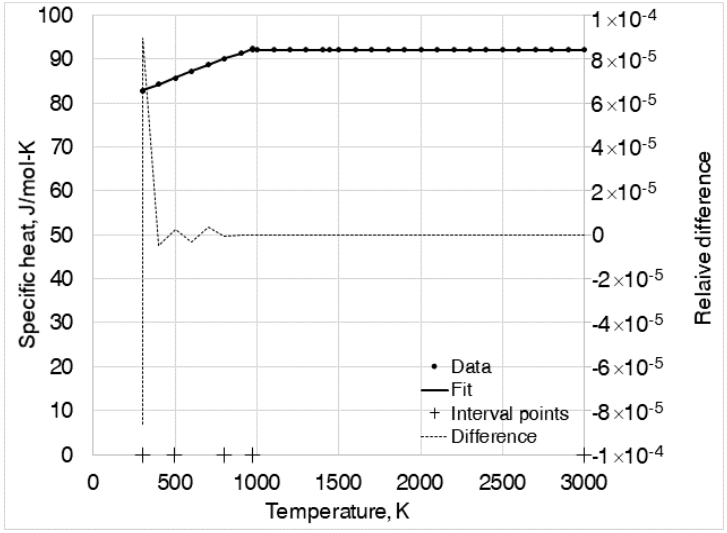


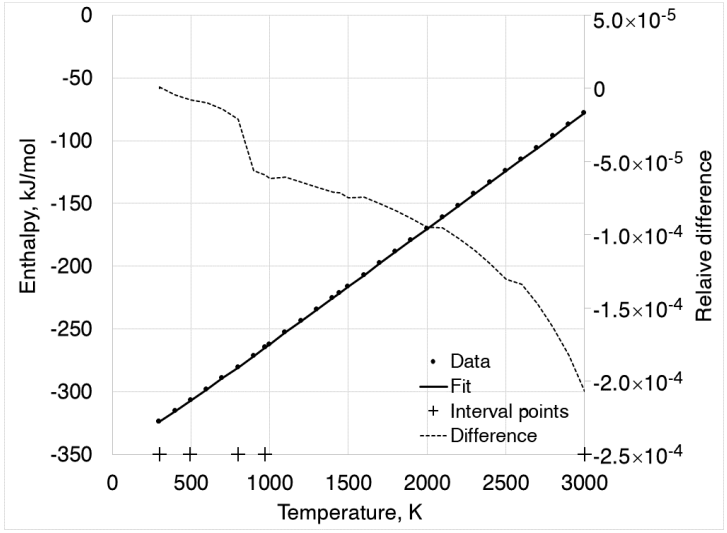


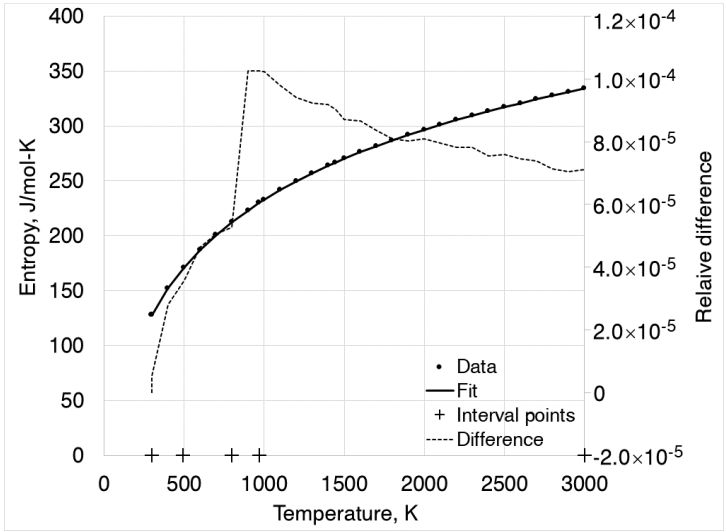


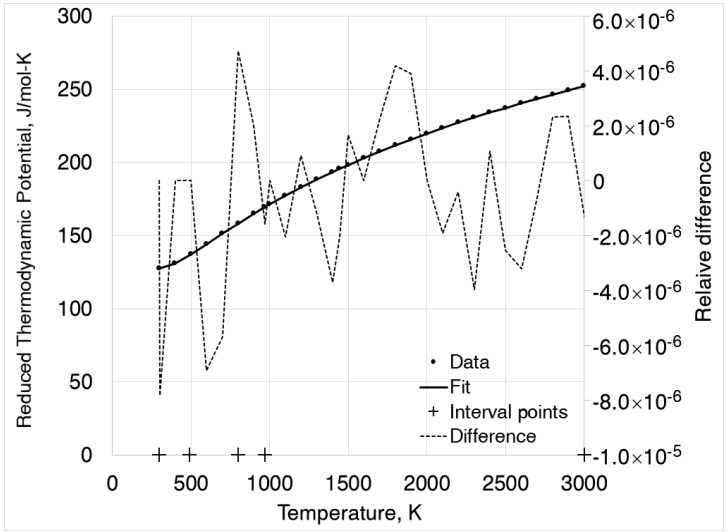
Sodium Sulfide (Na_2S) – Na2S1(l)

MW = 78.04 g/mol – $\Delta_f H^o$(298.15 K) = -323.94 kJ/mol – S^o(298.15 K) = 127.656 J/mol-K – p^o = 0.1 MPa



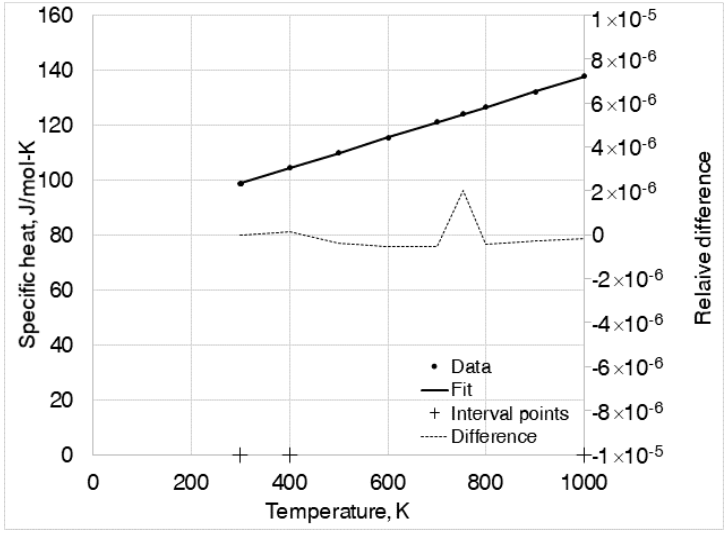


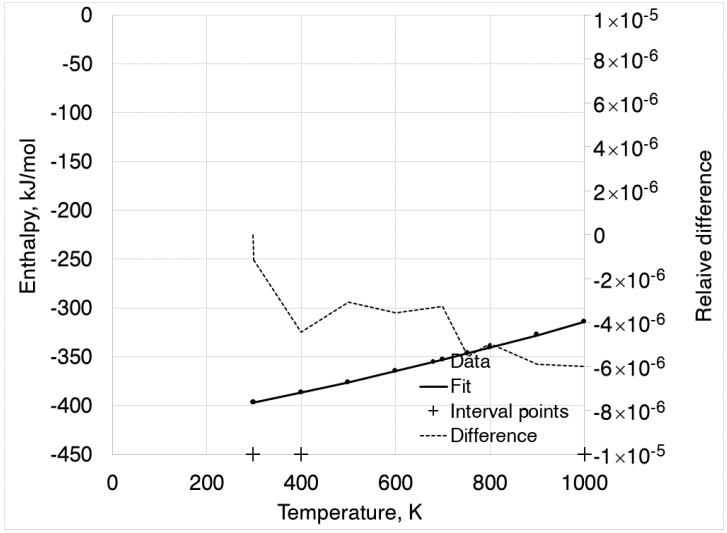


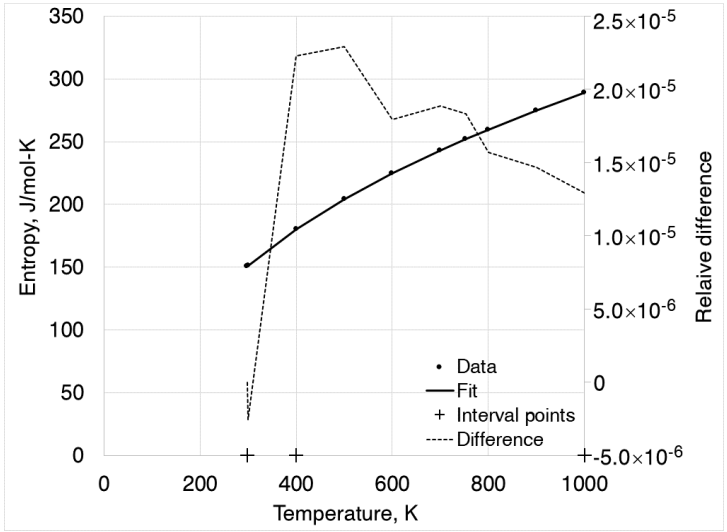


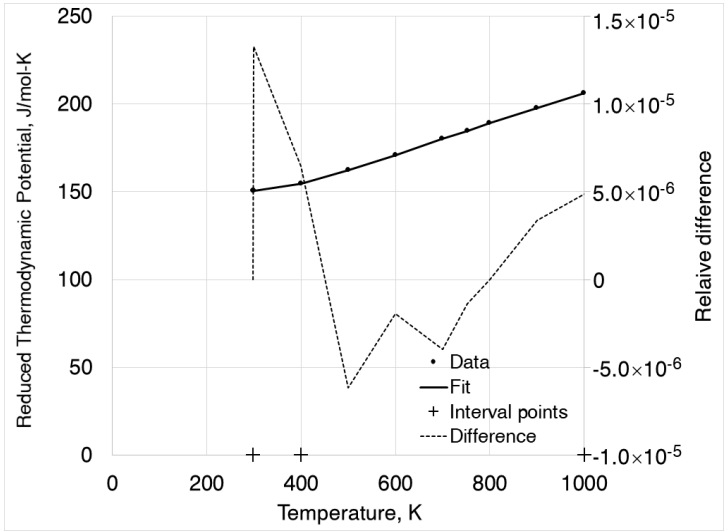
Sodium Sulfide, Beta (Na_2S_2) – Na2S2(cr)

MW = 110.1 g/mol – $\Delta_f H^o$(298.15 K) = -397.062 kJ/mol – S^o(298.15 K) = 150.624 J/mol-K – p^o = 0.1 MPa



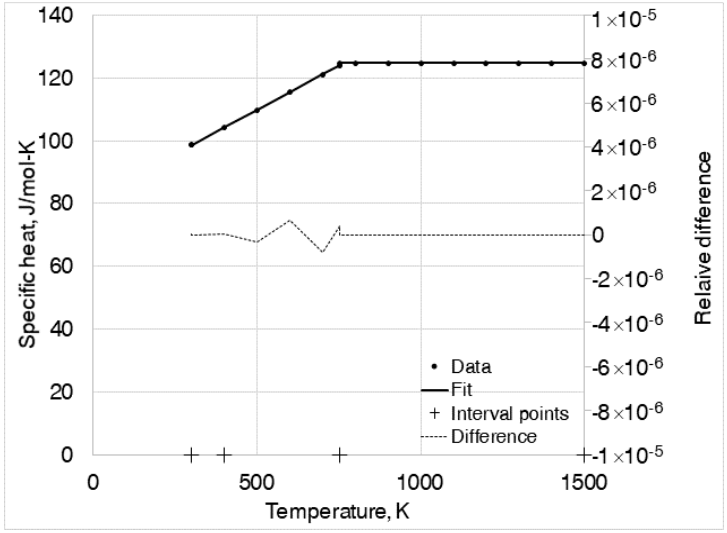


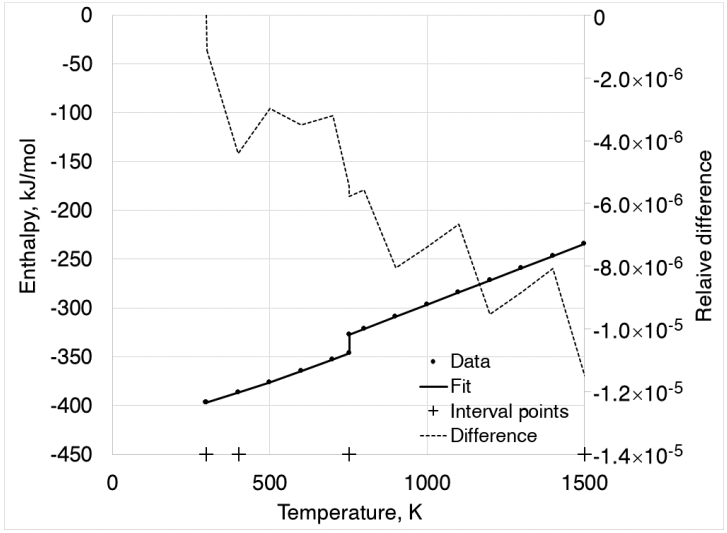


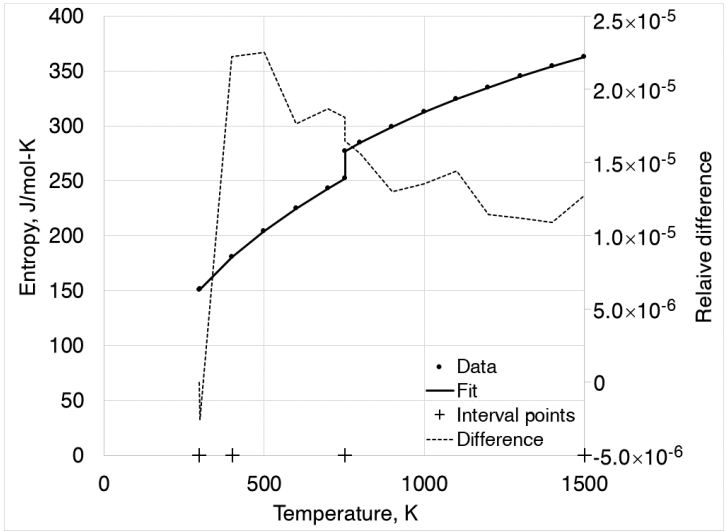


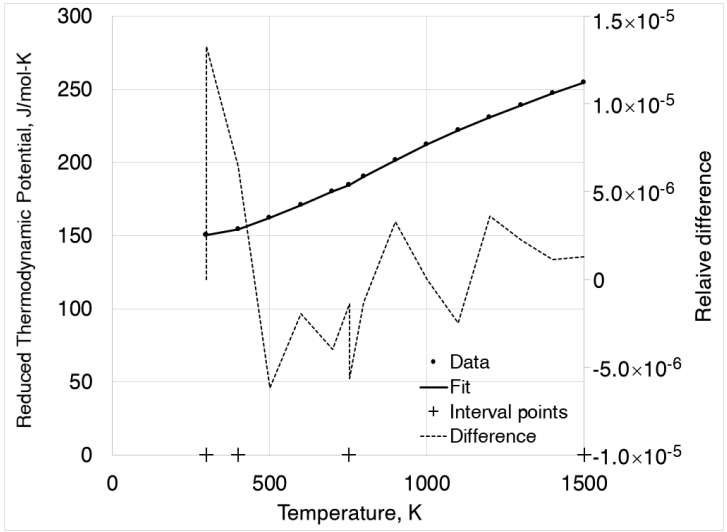
Sodium Sulfide (Na_2S_2) – Na2S2(cr,l)

MW = 110.1 g/mol – $\Delta_f H^o$(298.15 K) = -397.062 kJ/mol – S^o(298.15 K) = 150.624 J/mol-K – p^o = 0.1 MPa



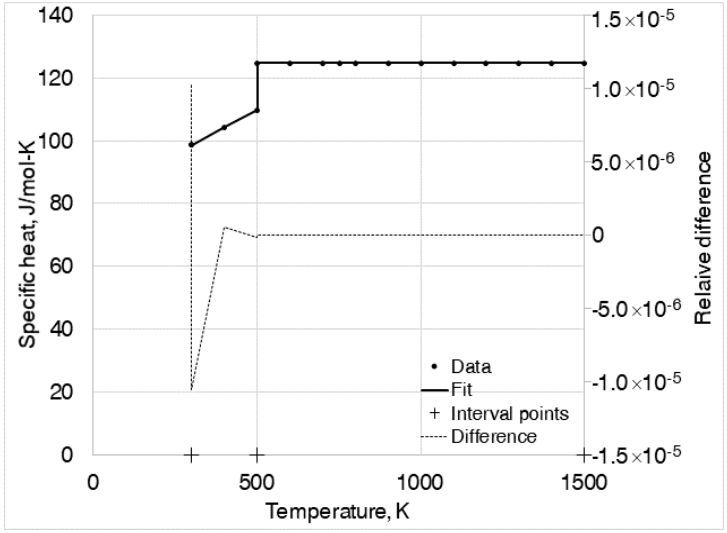


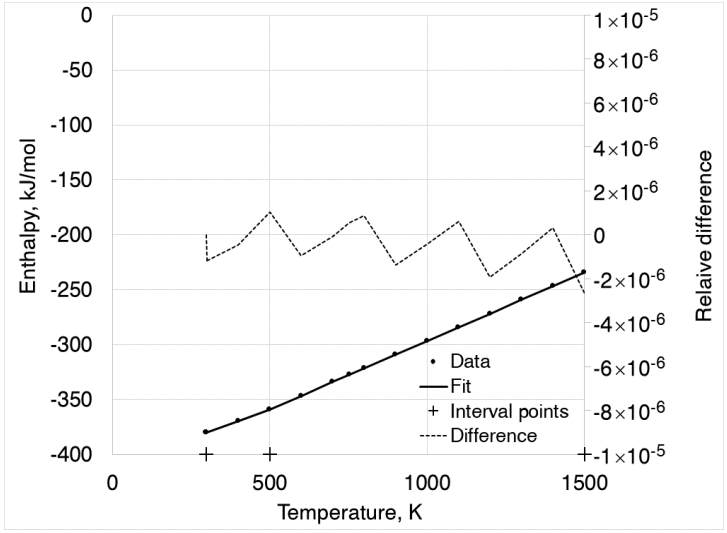


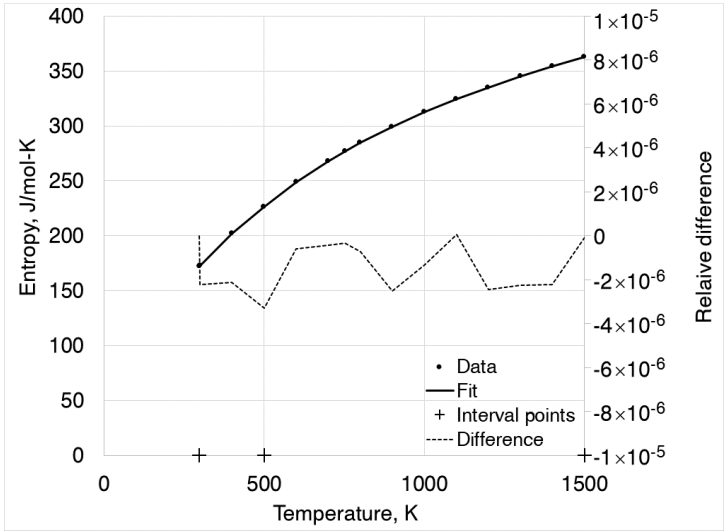


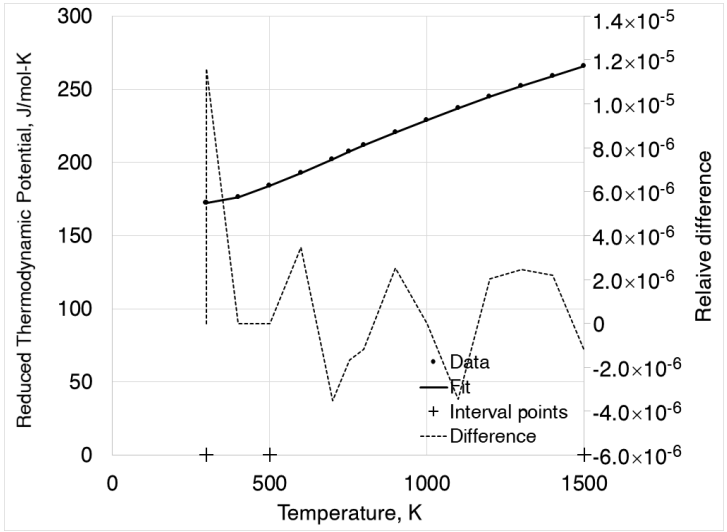
Sodium Sulfide (Na_2S_2) – Na2S2(l)

MW = 110.1 g/mol – $\Delta_f H^o$(298.15 K) = -380.114 kJ/mol – S^o(298.15 K) = 172.365 J/mol-K – p^o = 0.1 MPa



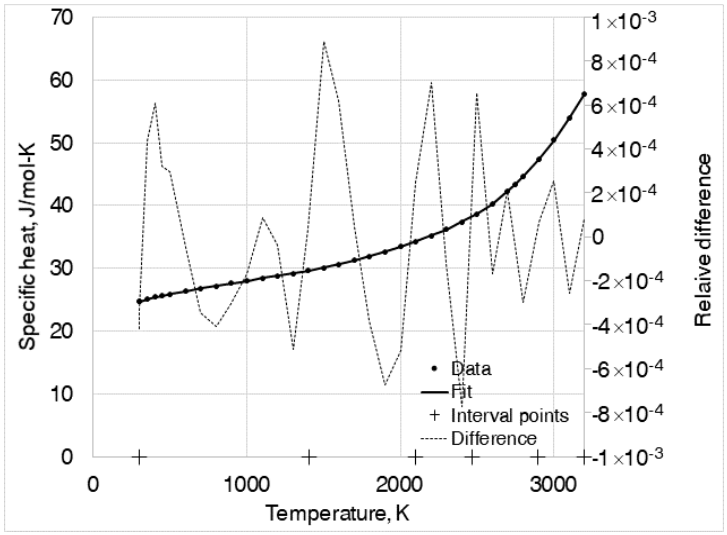


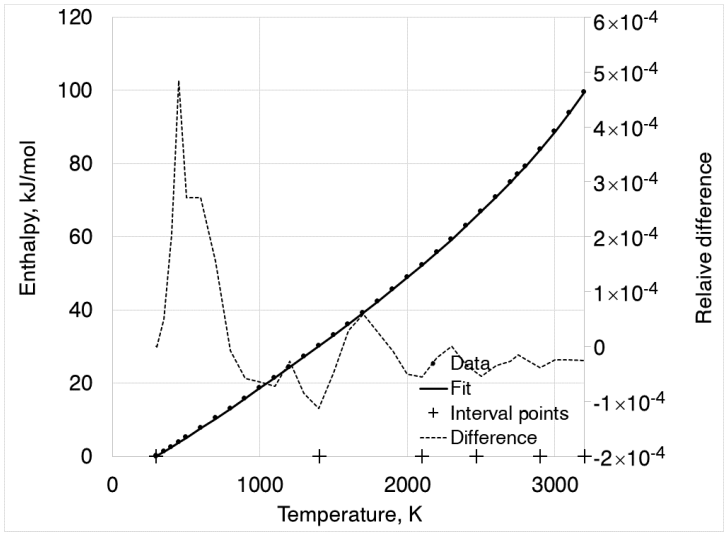


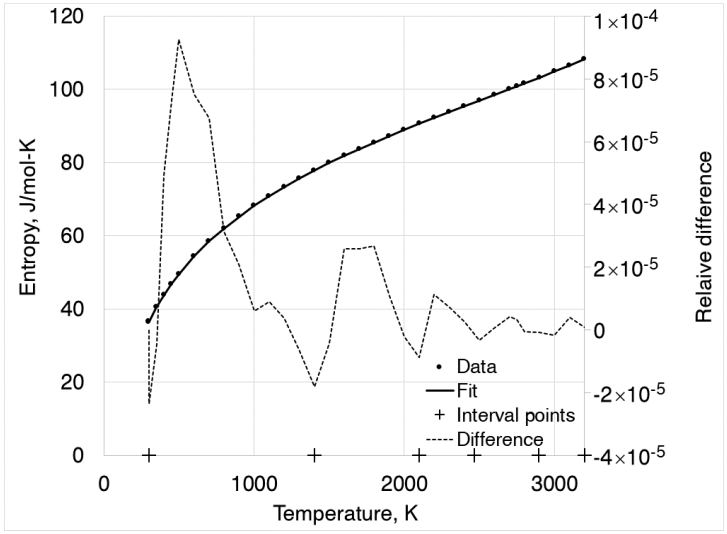


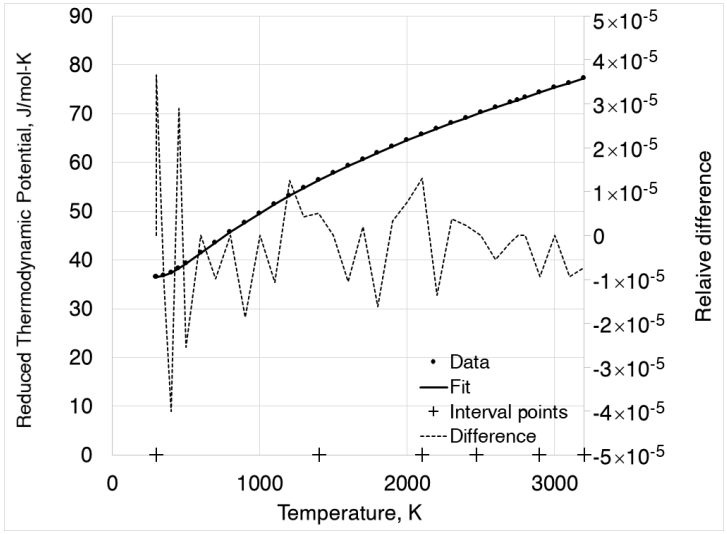
Niobium (Nb) – Nb1(cr)

MW = 92.9064 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 36.464 J/mol-K – p^o = 0.1 MPa



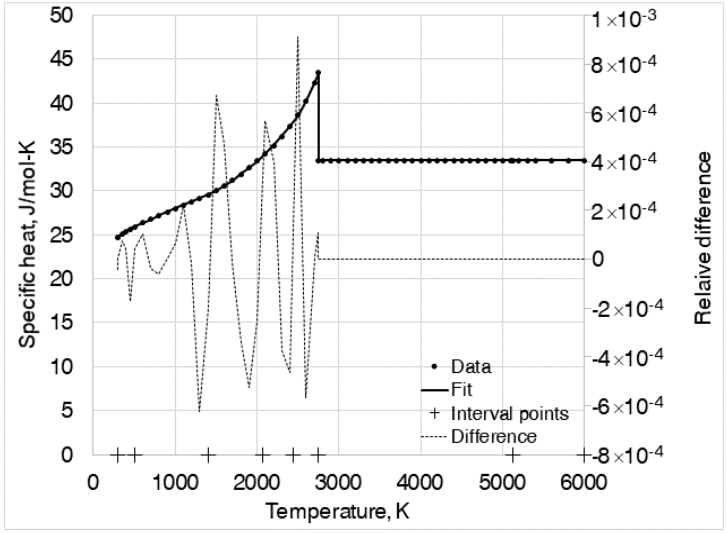


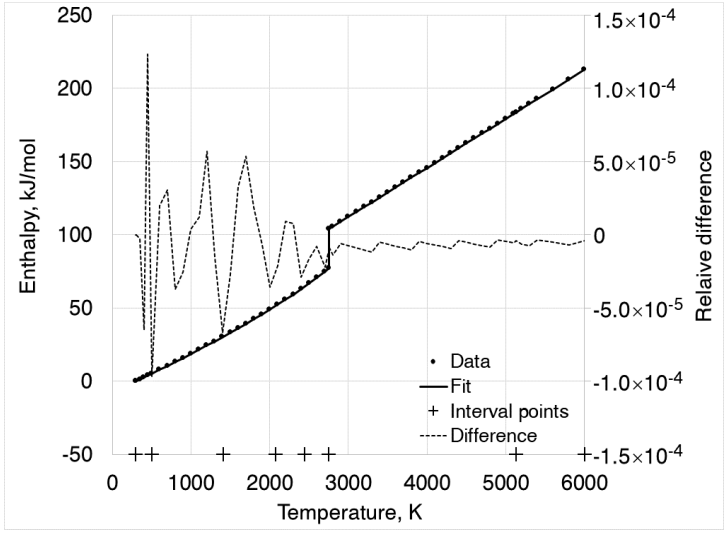


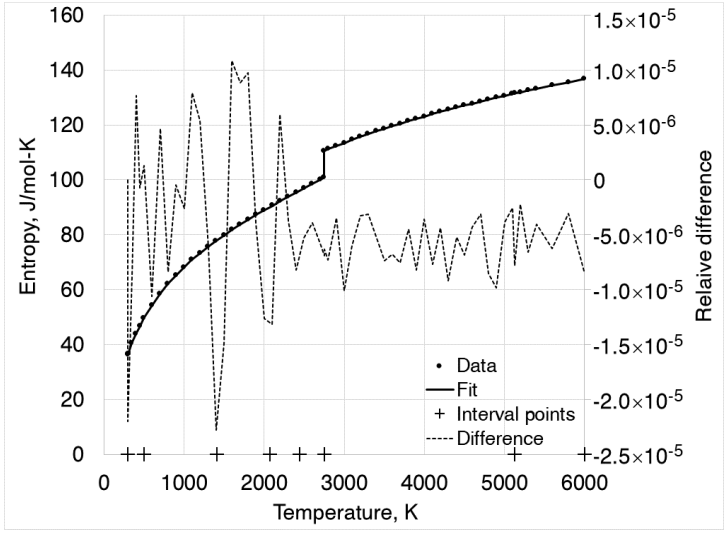


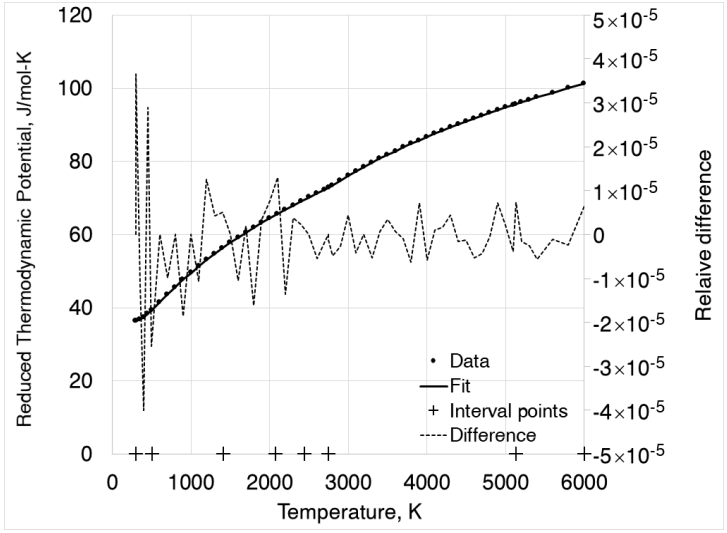
Niobium (Nb) – Nb1(cr,l)

MW = 92.9064 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 36.464 J/mol-K – p^o = 0.1 MPa



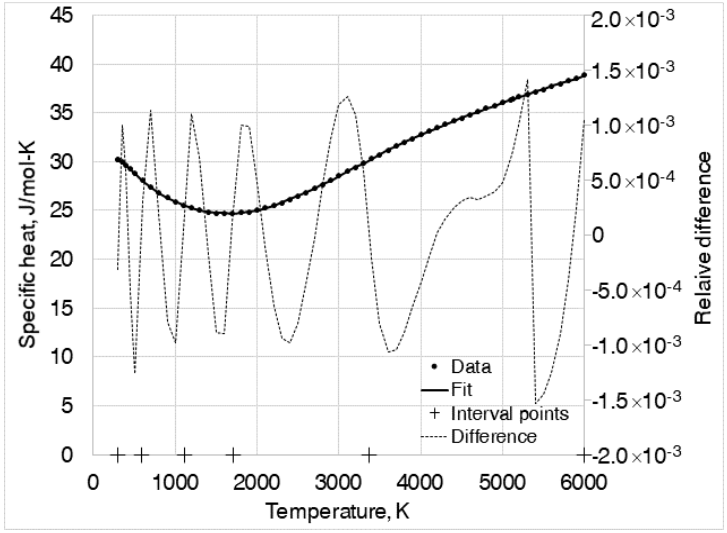


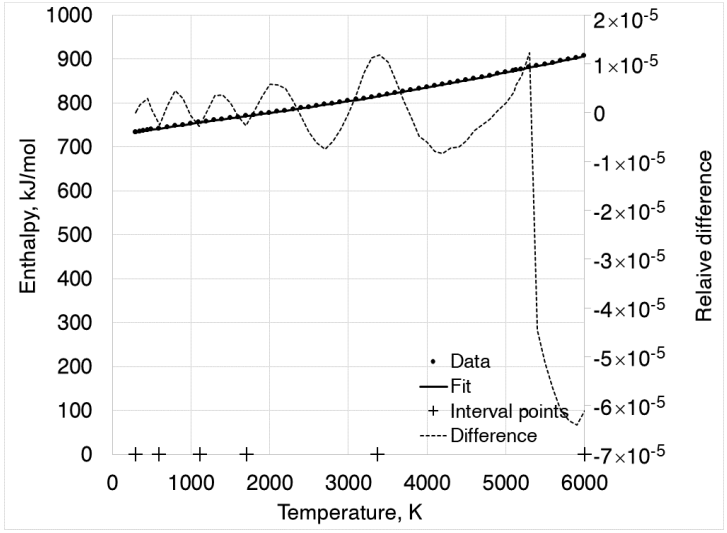


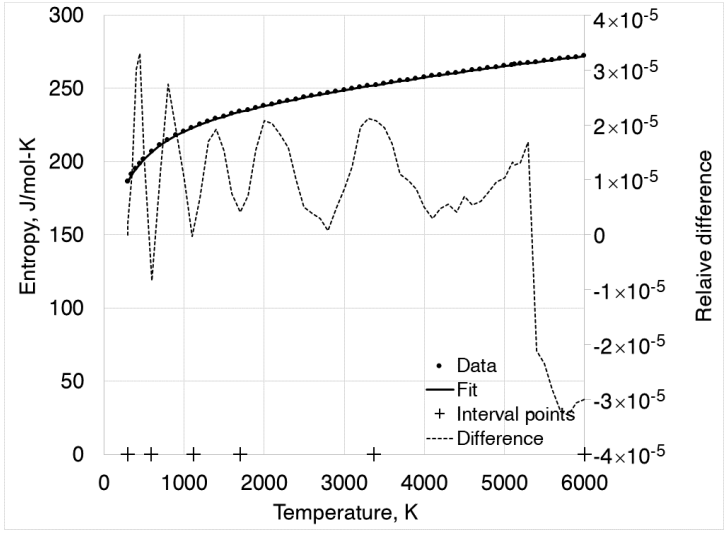


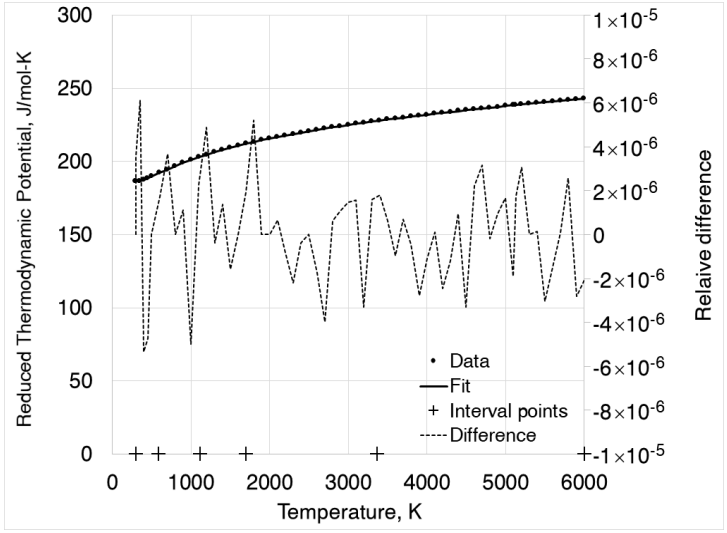
Niobium (Nb) – Nb1(g)

MW = 92.9064 g/mol – $\Delta_f H^o$(298.15 K) = 733.037 kJ/mol – S^o(298.15 K) = 186.259 J/mol-K – p^o = 0.1 MPa



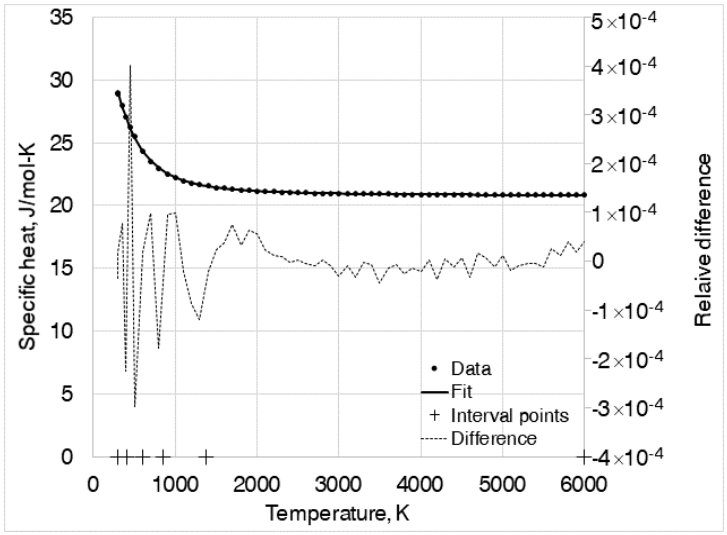


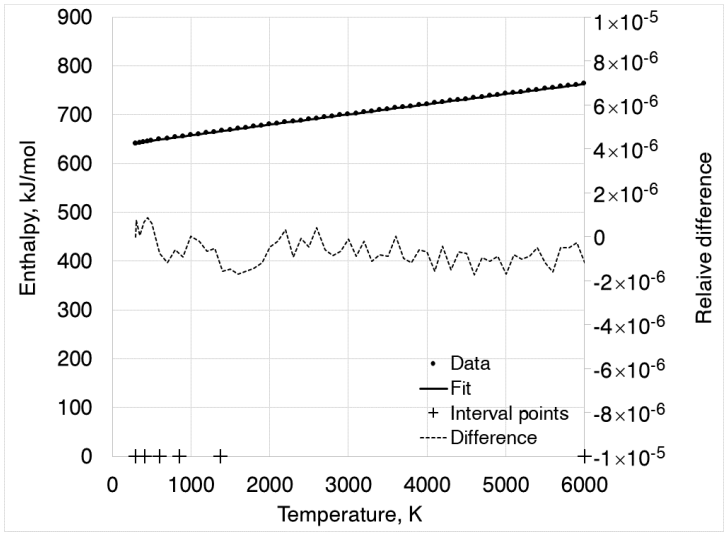


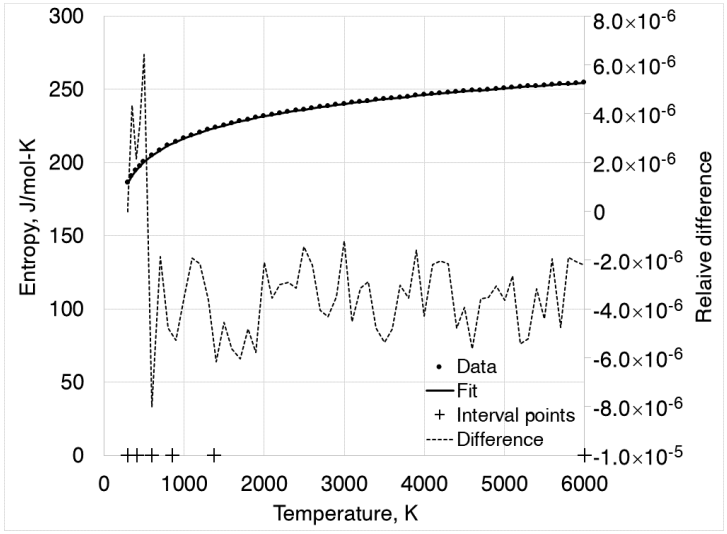


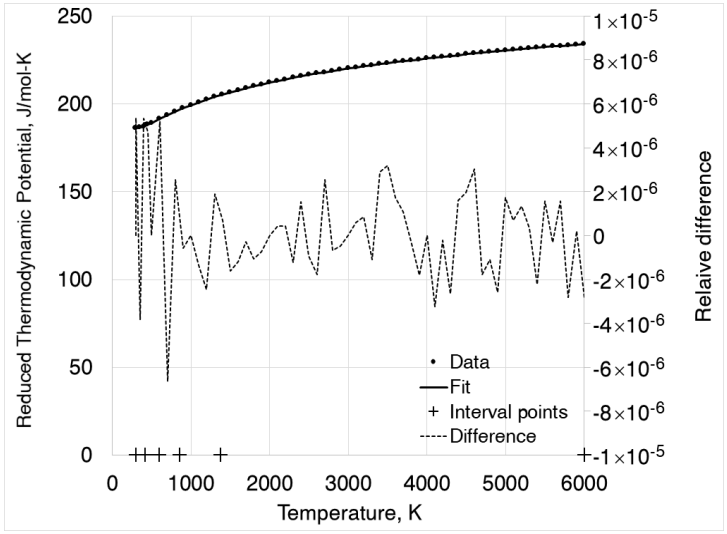
Niobium, Ion (Nb^-) – Nb1-(g)

MW = 92.9069 g/mol – $\Delta_f H^o$(298.15 K) = 640.978 kJ/mol – S^o(298.15 K) = 186.095 J/mol-K – p^o = 0.1 MPa



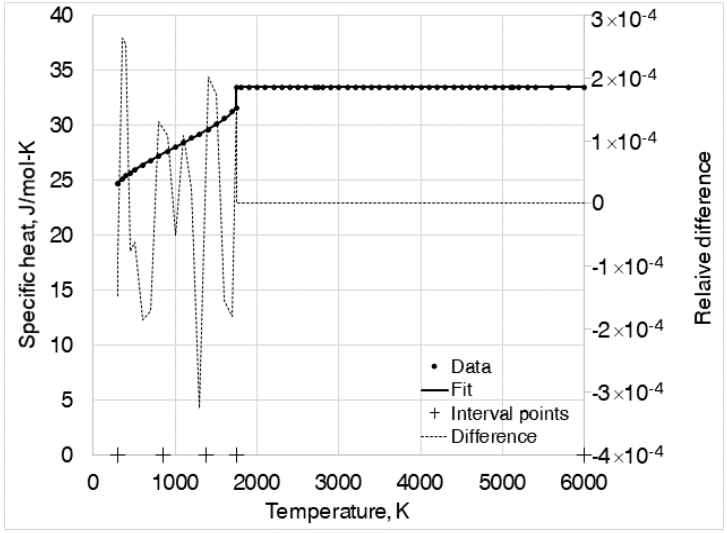


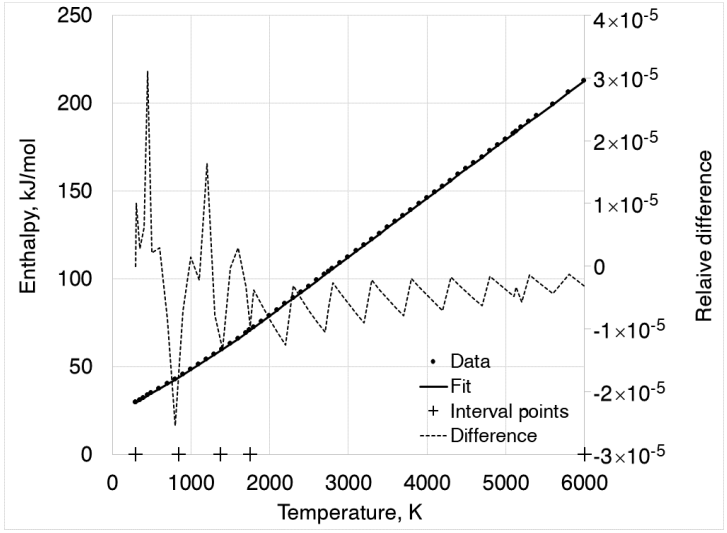


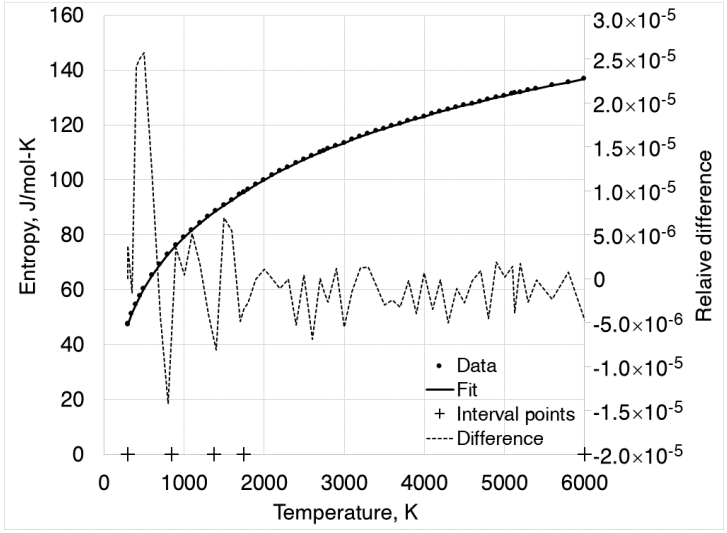


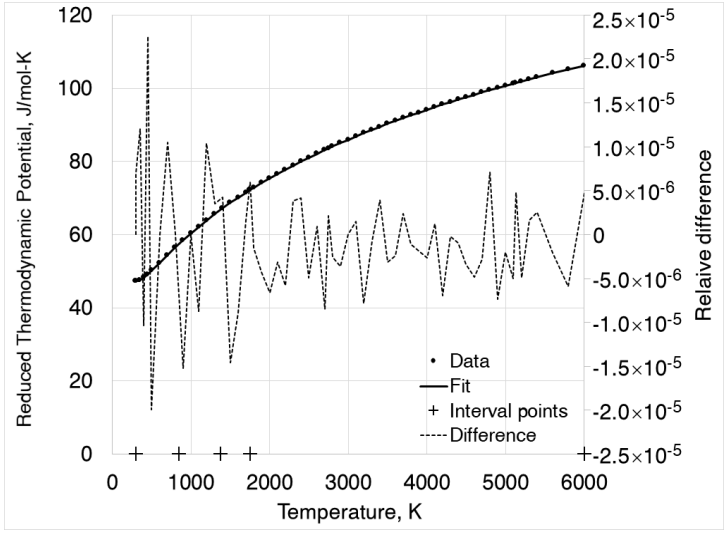
Niobium (Nb) – Nb1(l)

MW = 92.9064 g/mol – $\Delta_f H^o$(298.15 K) = 29.647 kJ/mol – S^o(298.15 K) = 47.302 J/mol-K – p^o = 0.1 MPa



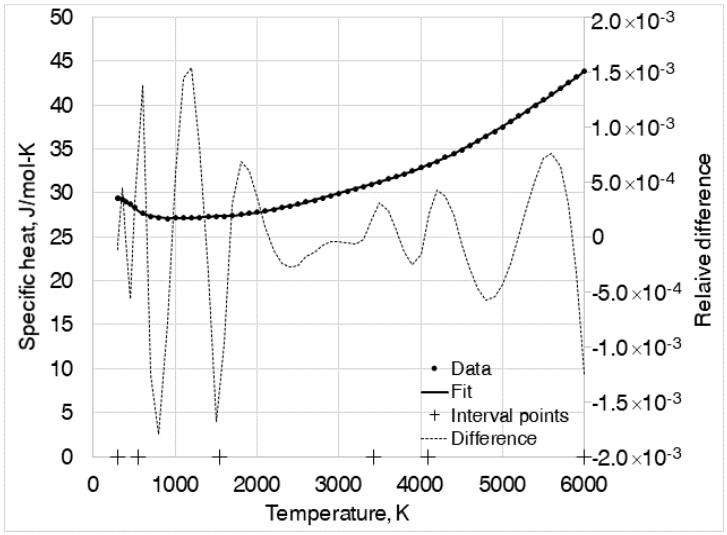


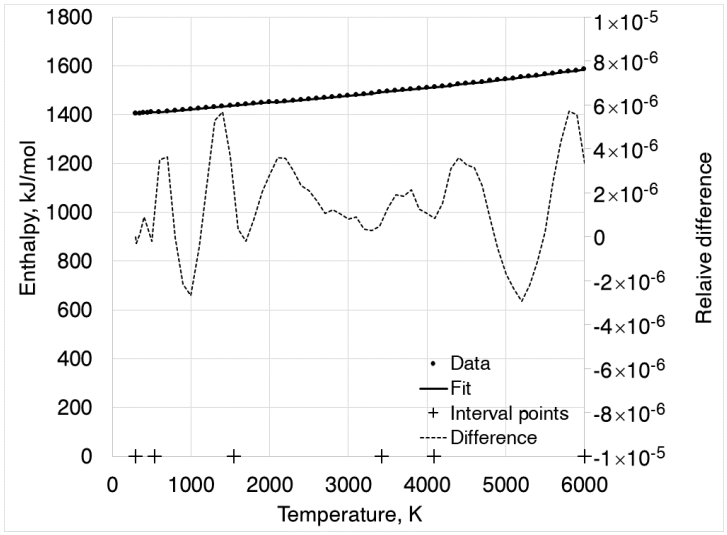


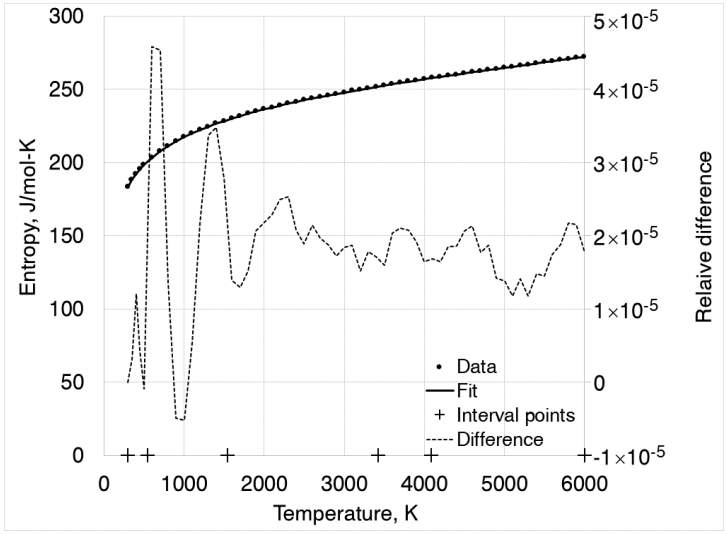


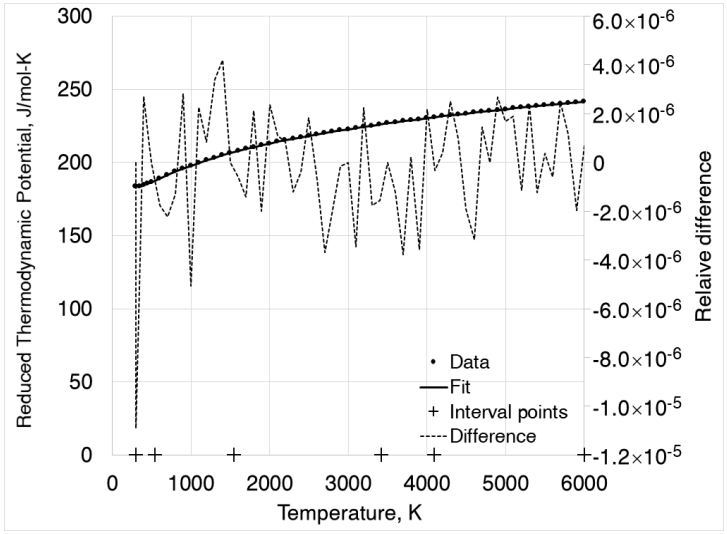
Niobium, Ion (Nb^+) – Nb1+(g)

MW = 92.9059 g/mol – $\Delta_f H^o$(298.15 K) = 1402.954 kJ/mol – S^o(298.15 K) = 183.333 J/mol-K – p^o = 0.1 MPa



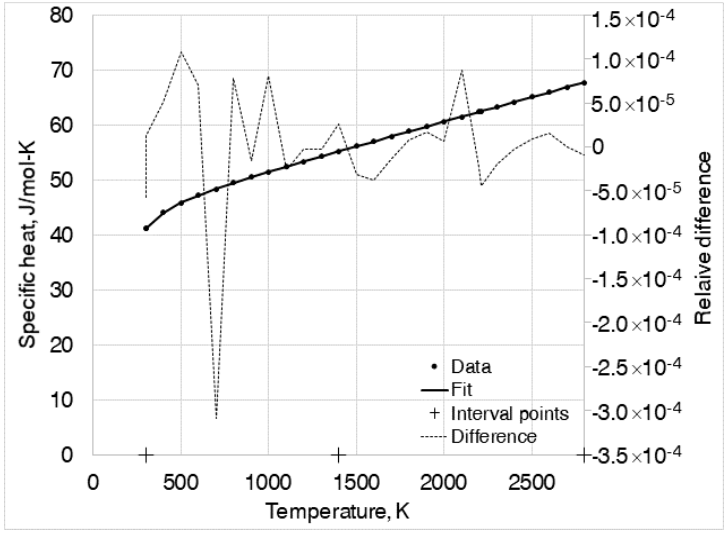


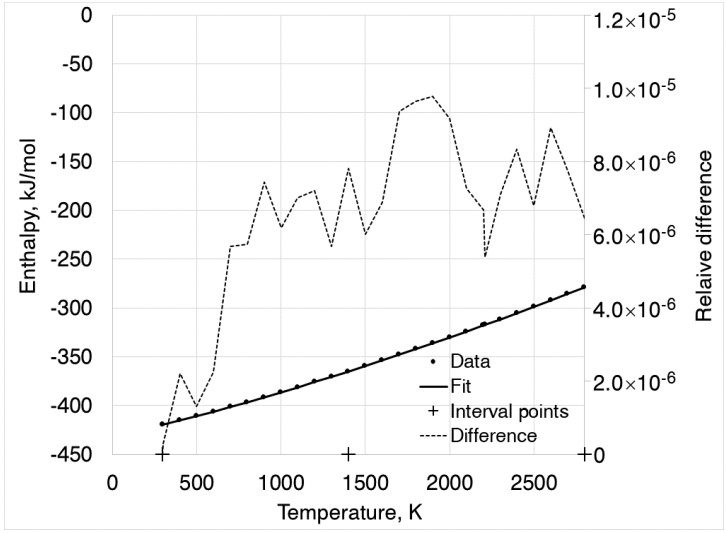


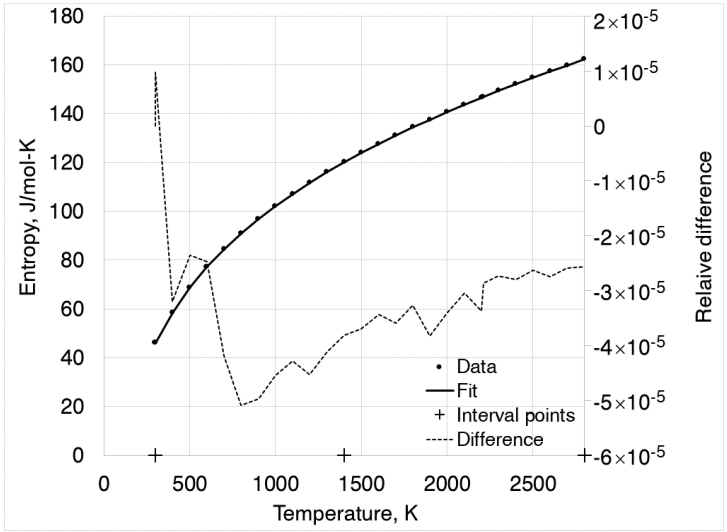


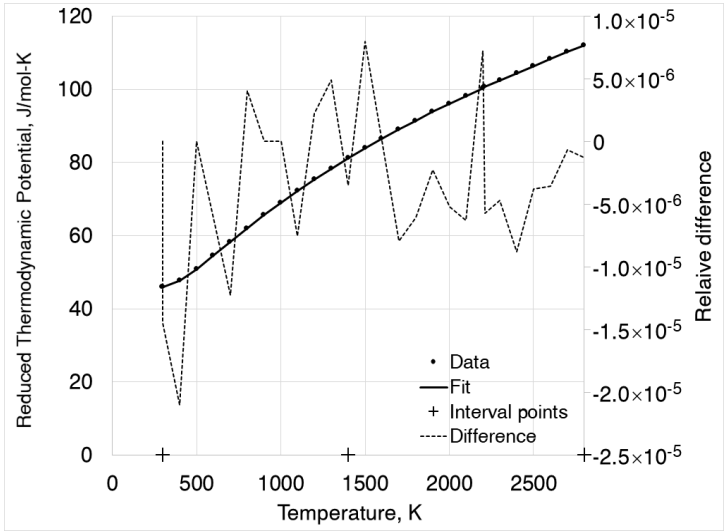
Niobium Oxide (NbO) – Nb1O1(cr)

MW = 108.9058 g/mol – $\Delta_f H^o$(298.15 K) = -419.655 kJ/mol – S^o(298.15 K) = 46.024 J/mol-K – p^o = 0.1 MPa



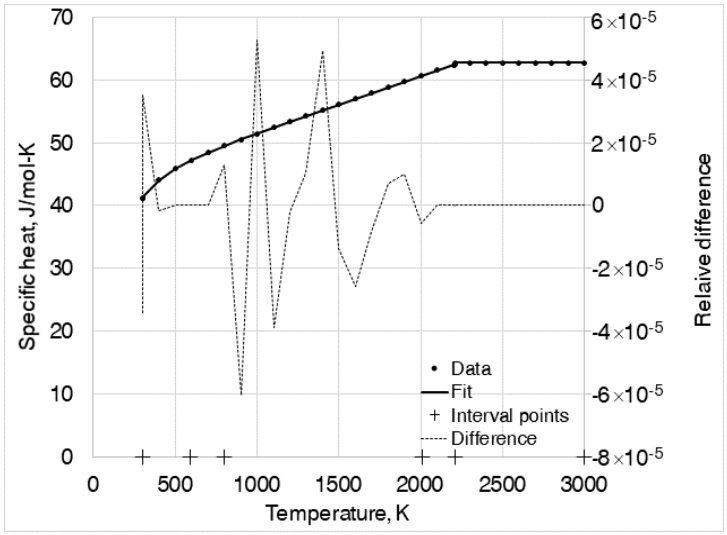


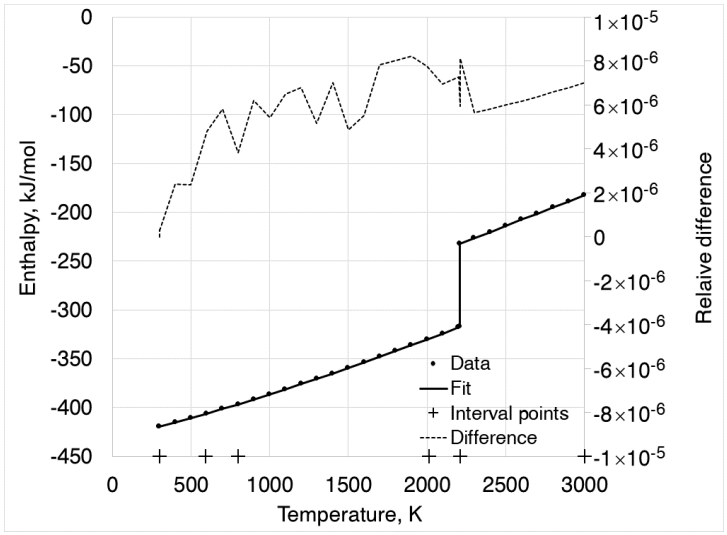


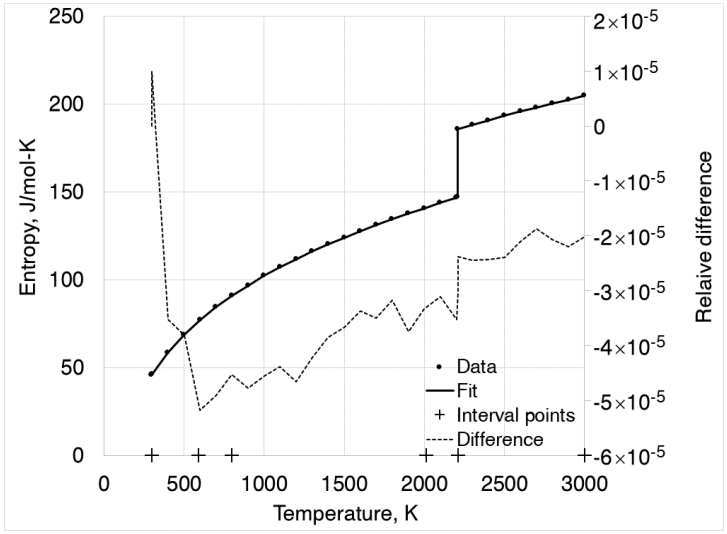


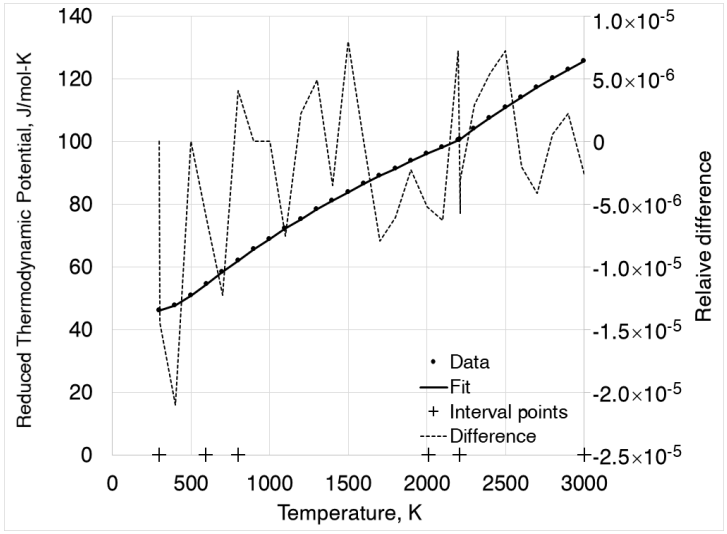
Niobium Oxide (NbO) – Nb1O1(cr,l)

MW = 108.9058 g/mol – $\Delta_f H^o$(298.15 K) = -419.655 kJ/mol – S^o(298.15 K) = 46.024 J/mol-K – p^o = 0.1 MPa



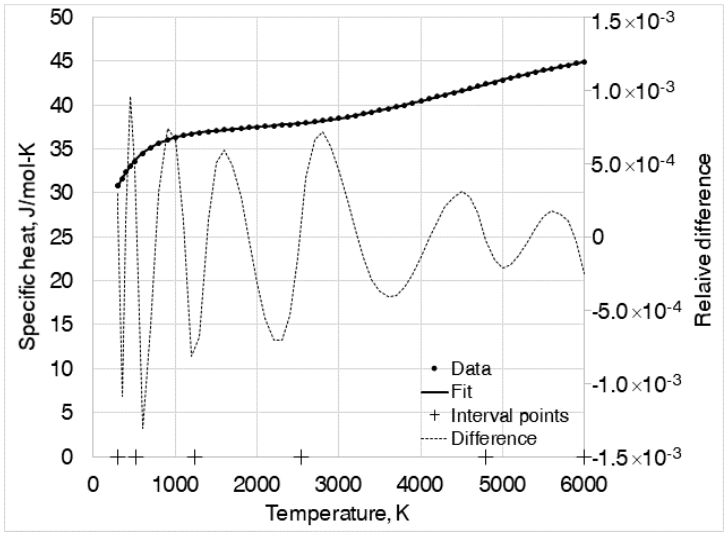


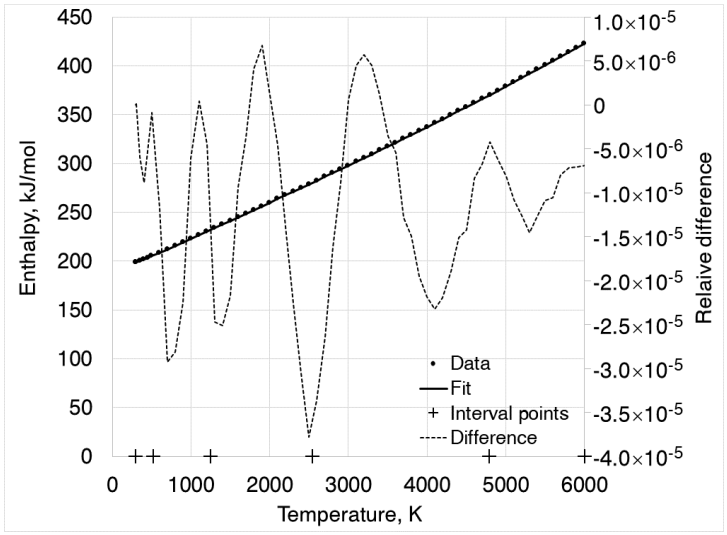


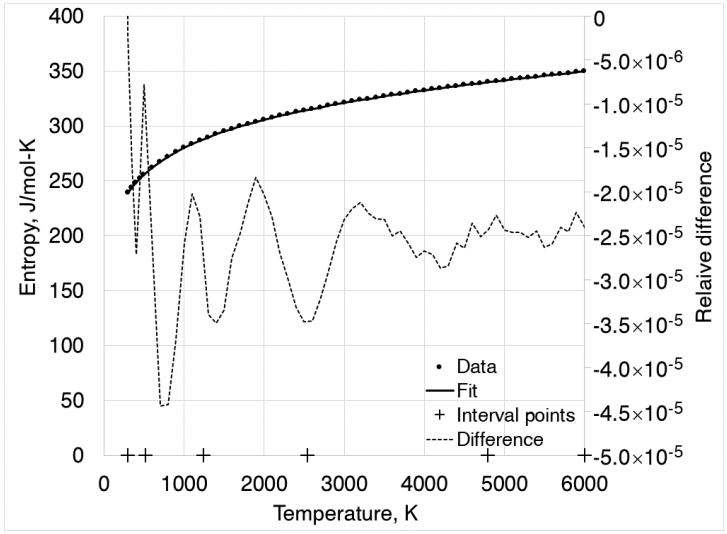


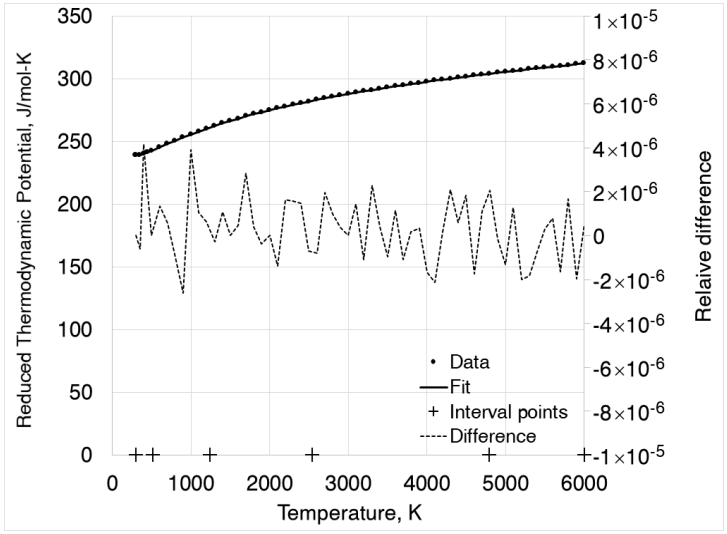
Niobium Oxide (NbO) – Nb1O1(g)

MW = 108.9058 g/mol – $\Delta_f H^o$(298.15 K) = 198.74 kJ/mol – S^o(298.15 K) = 238.976 J/mol-K – p^o = 0.1 MPa



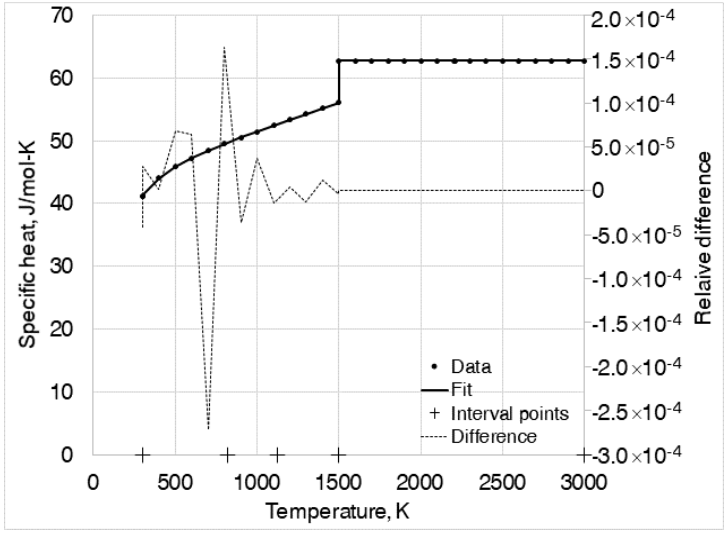


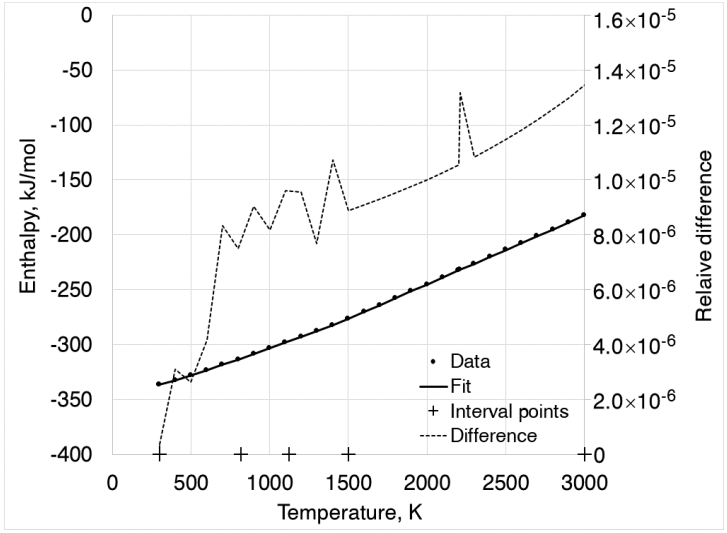


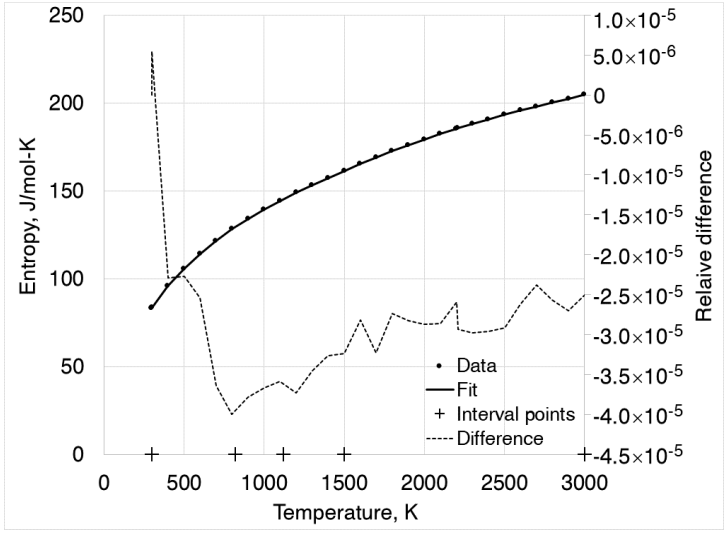


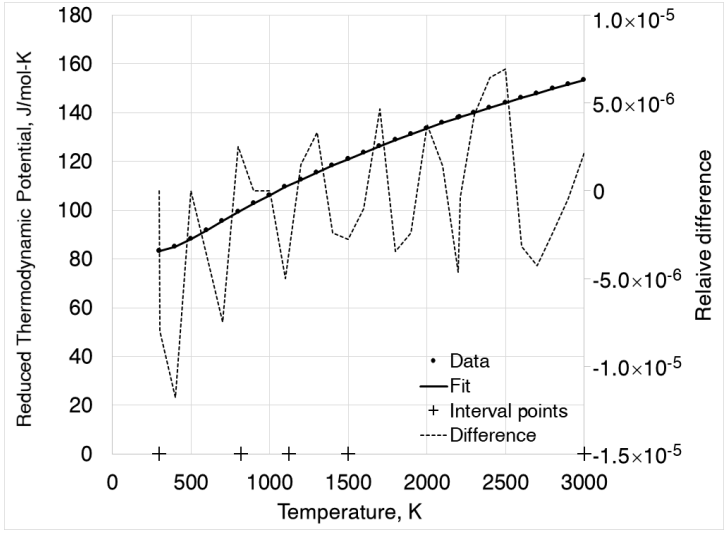
Niobium Oxide (NbO) – Nb1O1(l)

MW = 108.9058 g/mol – $\Delta_f H^o$(298.15 K) = -336.743 kJ/mol – S^o(298.15 K) = 83.233 J/mol-K – p^o = 0.1 MPa



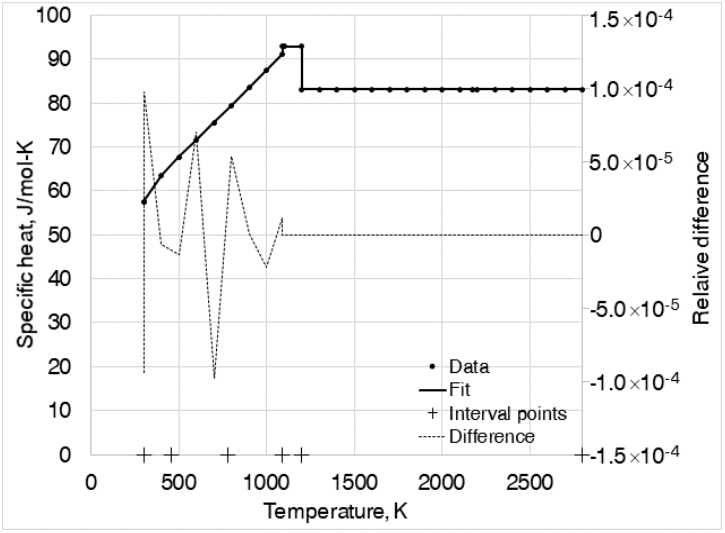


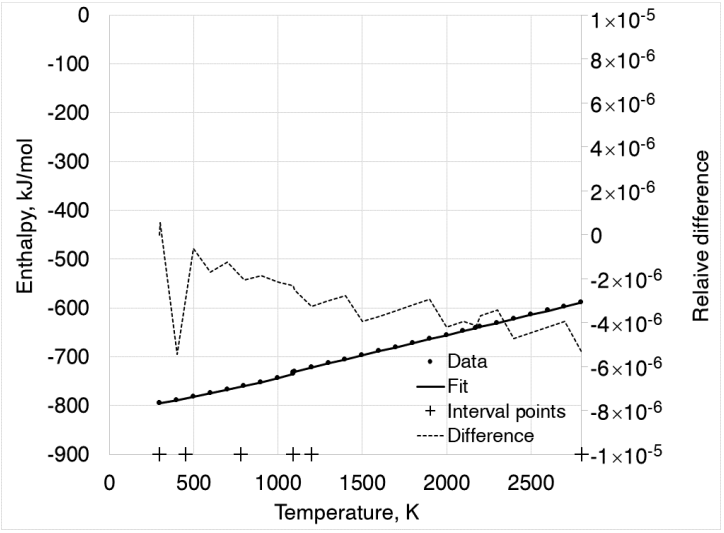


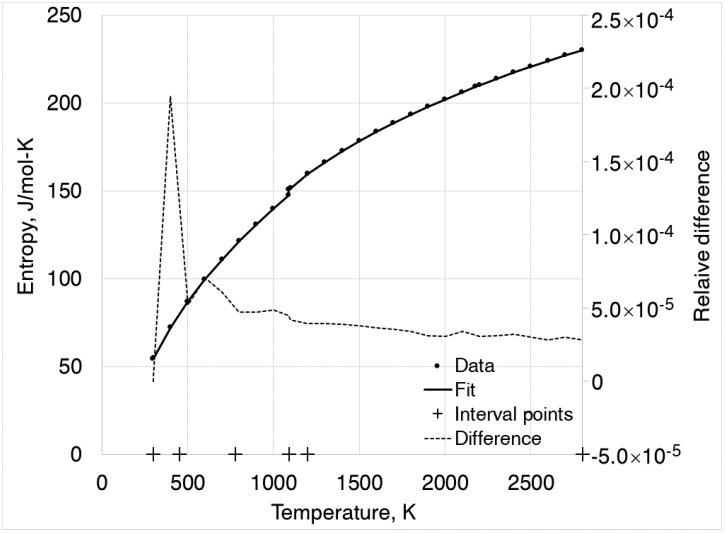


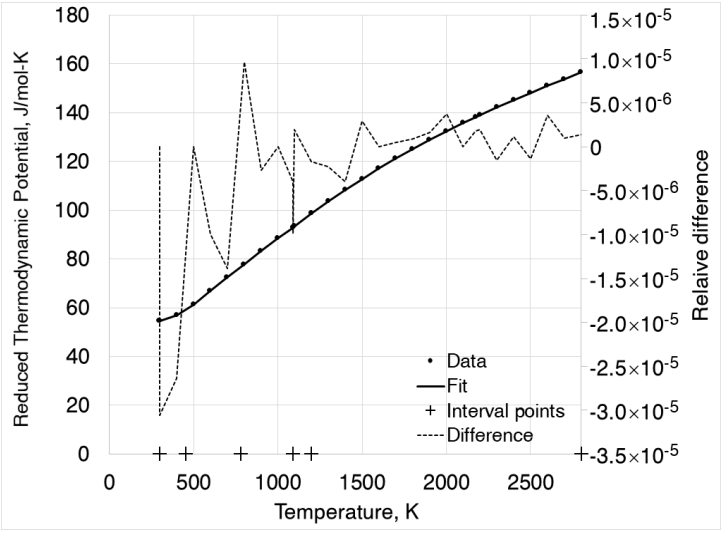
Niobium Oxide (NbO_2) – Nb1O2(cr)

MW = 124.9052 g/mol – $\Delta_f H^o$(298.15 K) = -794.96 kJ/mol – S^o(298.15 K) = 54.506 J/mol-K – p^o = 0.1 MPa



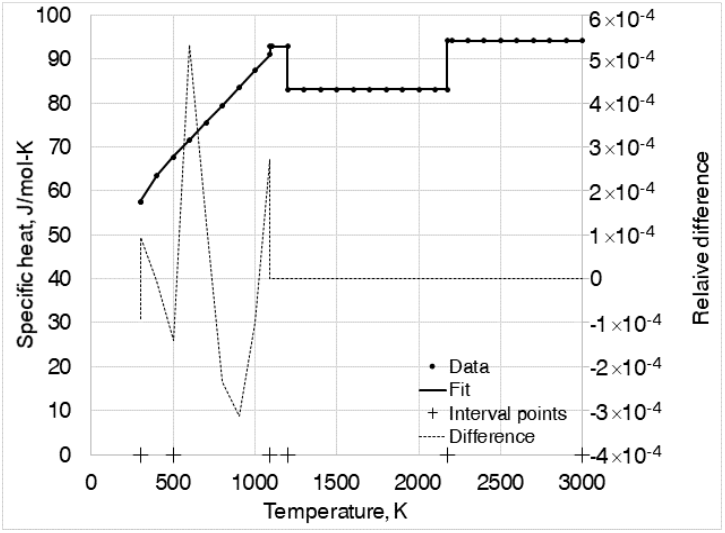


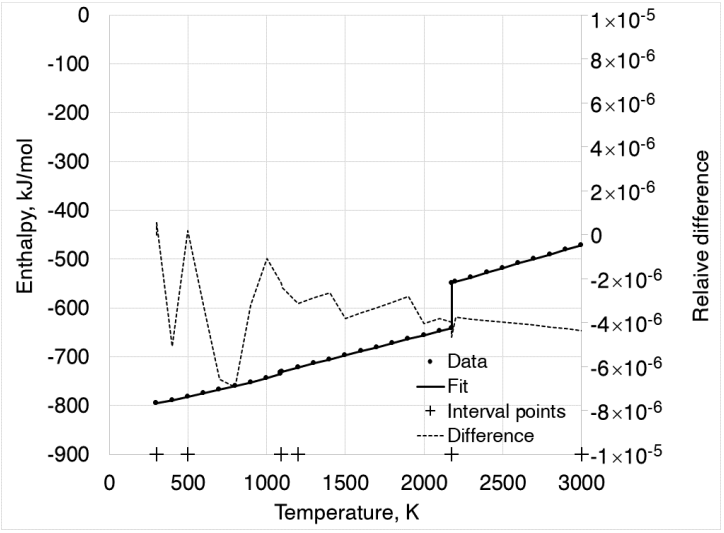


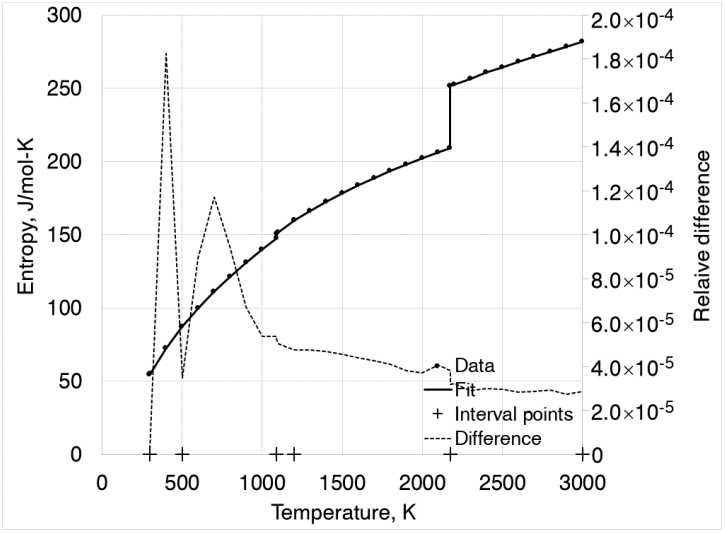


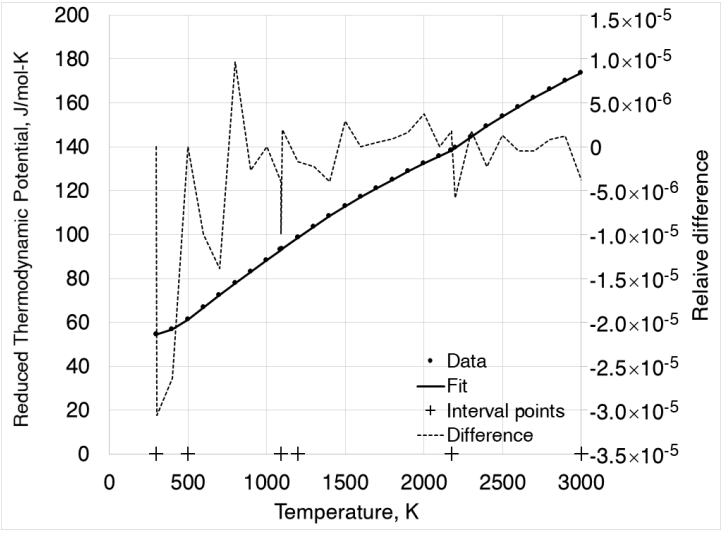
Niobium Oxide (NbO_2) – Nb1O2(cr,l)

MW = 124.9052 g/mol – $\Delta_f H^o$(298.15 K) = -794.96 kJ/mol – S^o(298.15 K) = 54.506 J/mol-K – p^o = 0.1 MPa



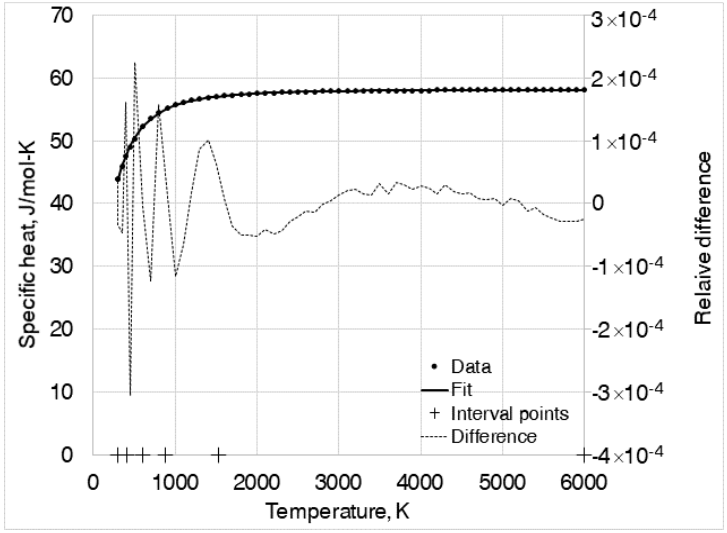


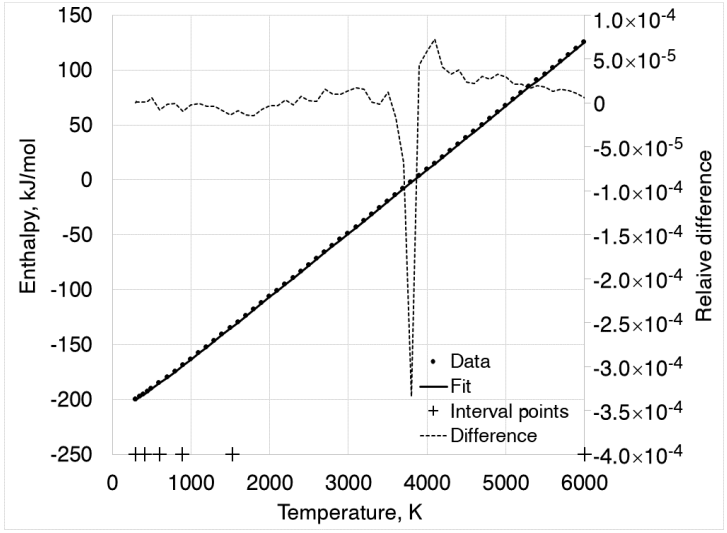


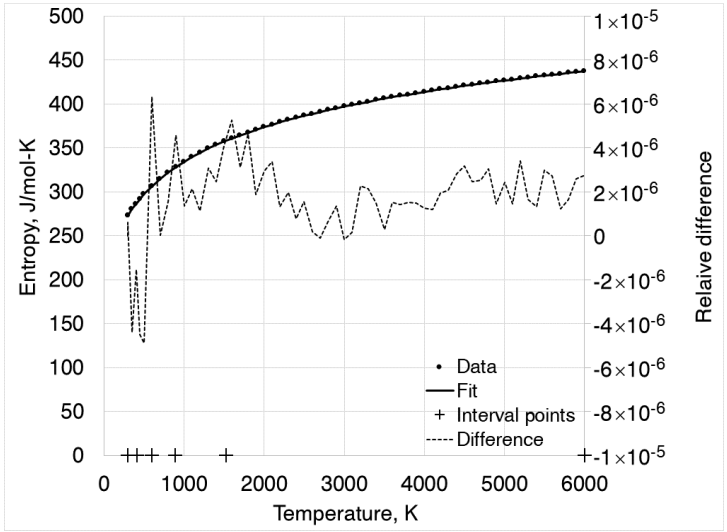


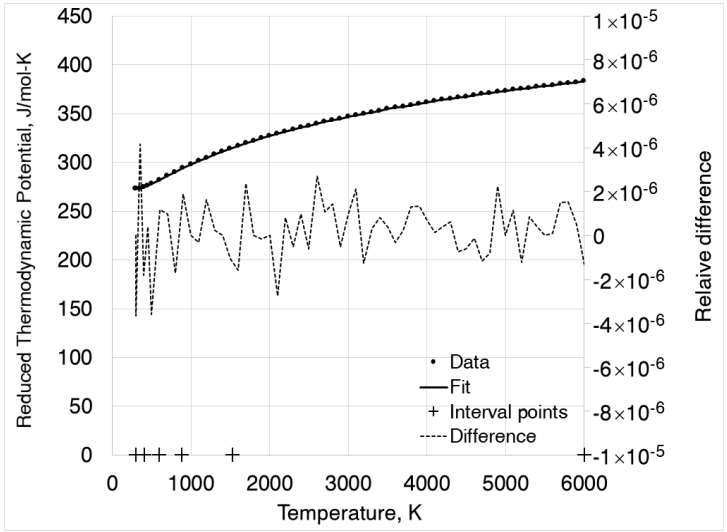
Niobium Oxide (NbO_2) – Nb1O2(g)

MW = 124.9052 g/mol – $\Delta_f H^o$(298.15 K) = -199.995 kJ/mol – S^o(298.15 K) = 272.921 J/mol-K – p^o = 0.1 MPa



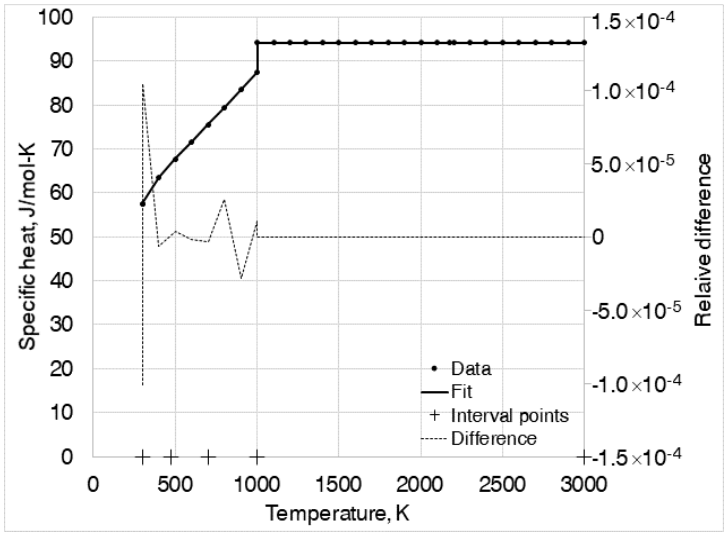


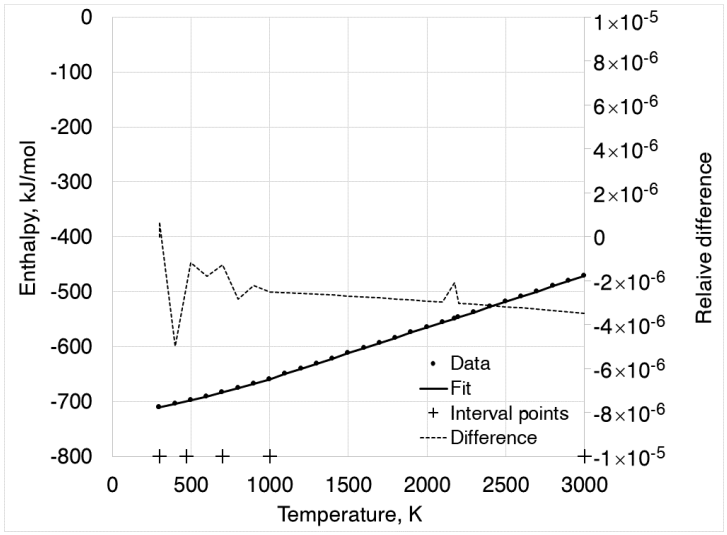


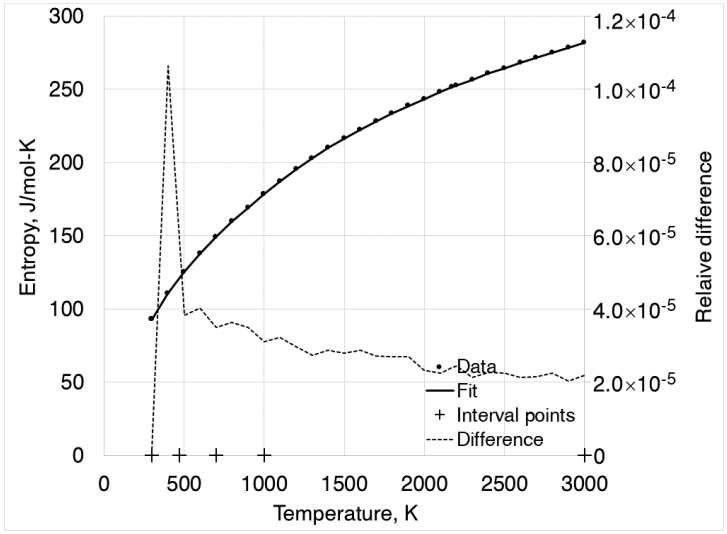


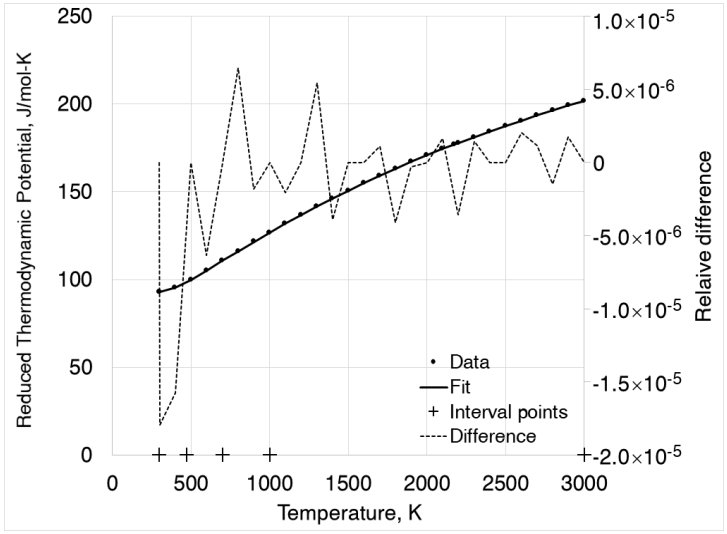
Niobium Oxide (NbO_2) – Nb1O2(l)

MW = 124.9052 g/mol – $\Delta_f H^o$(298.15 K) = -710.872 kJ/mol – S^o(298.15 K) = 92.835 J/mol-K – p^o = 0.1 MPa



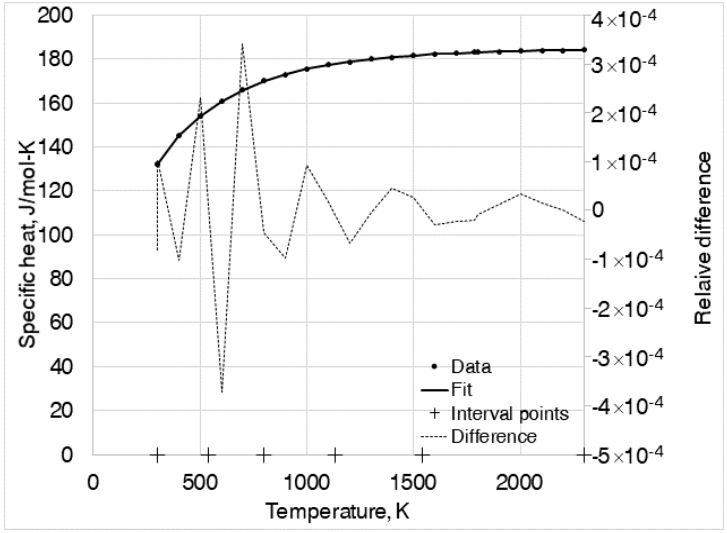


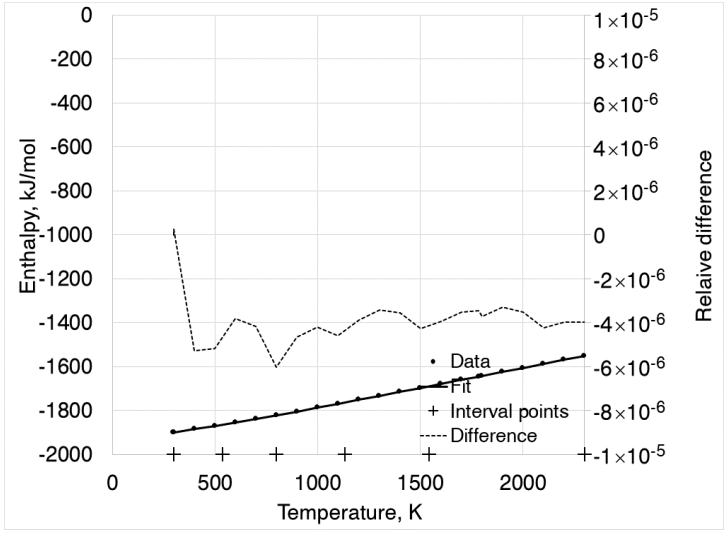


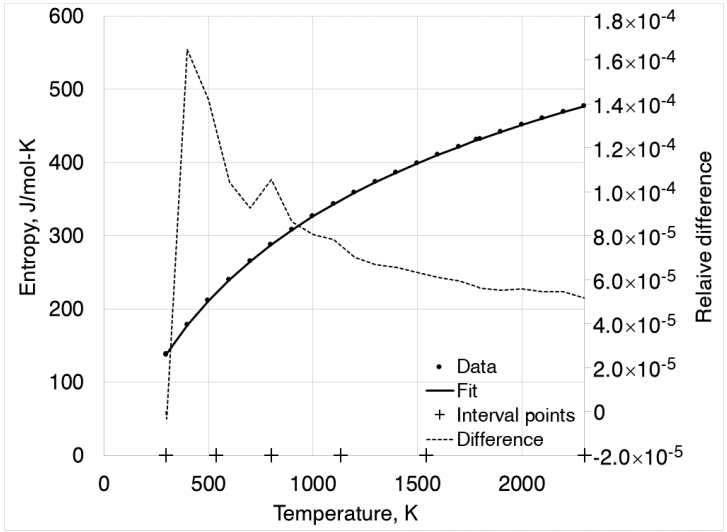


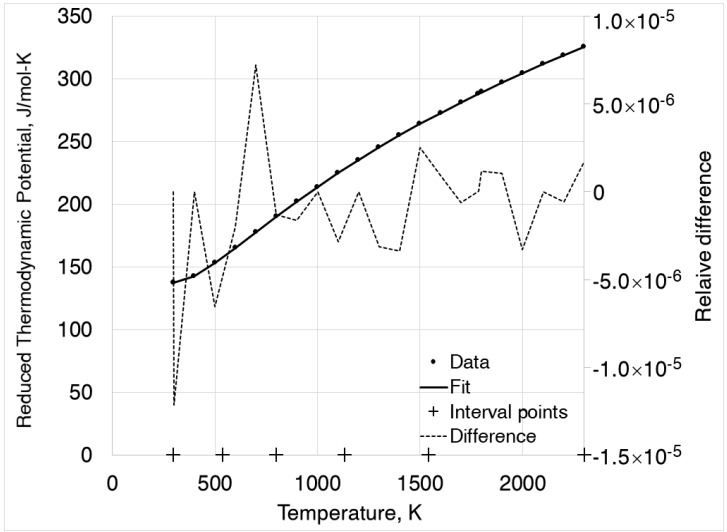
Niobium Oxide (Nb_2O_5) – Nb2O5(cr)

MW = 265.8098 g/mol – $\Delta_f H^o$(298.15 K) = -1899.536 kJ/mol – S^o(298.15 K) = 137.298 J/mol-K – p^o = 0.1 MPa



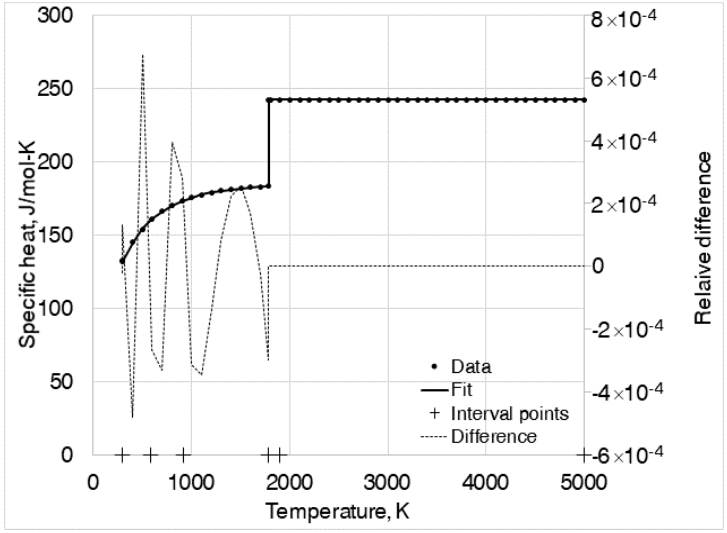


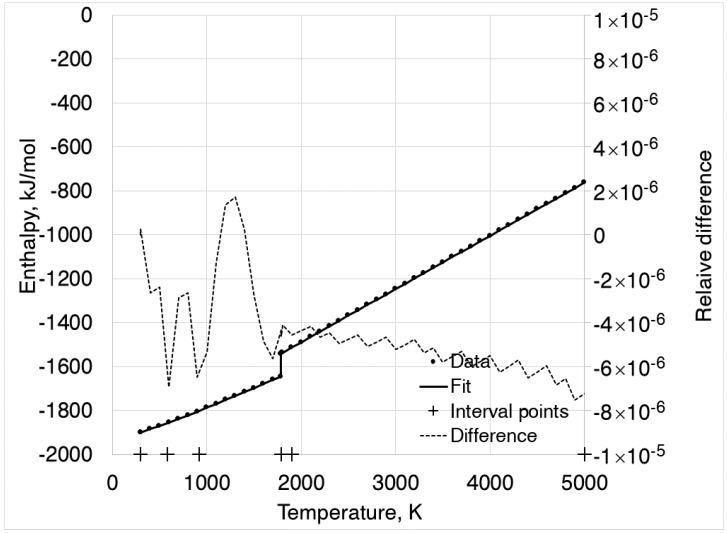


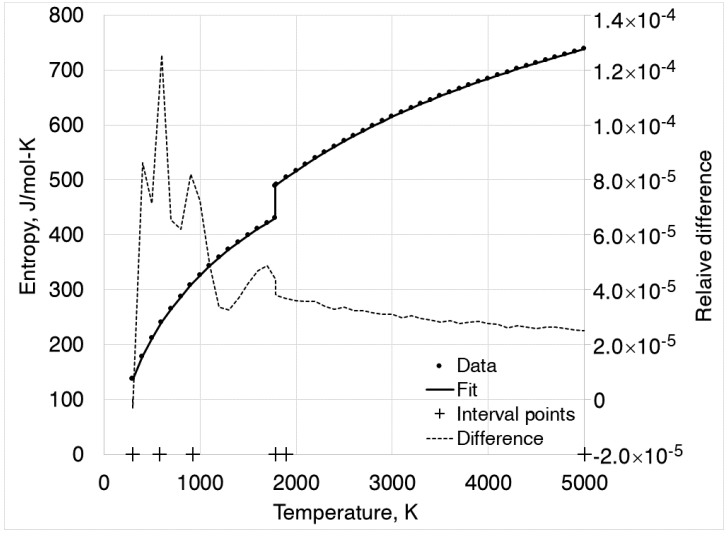


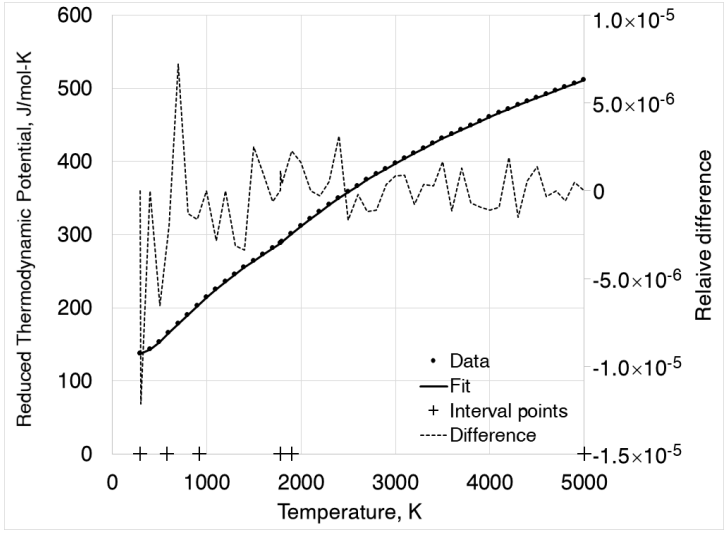
Niobium Oxide (Nb_2O_5) – Nb2O5(cr,l)

MW = 265.8098 g/mol – $\Delta_f H^o$(298.15 K) = -1899.536 kJ/mol – S^o(298.15 K) = 137.298 J/mol-K – p^o = 0.1 MPa



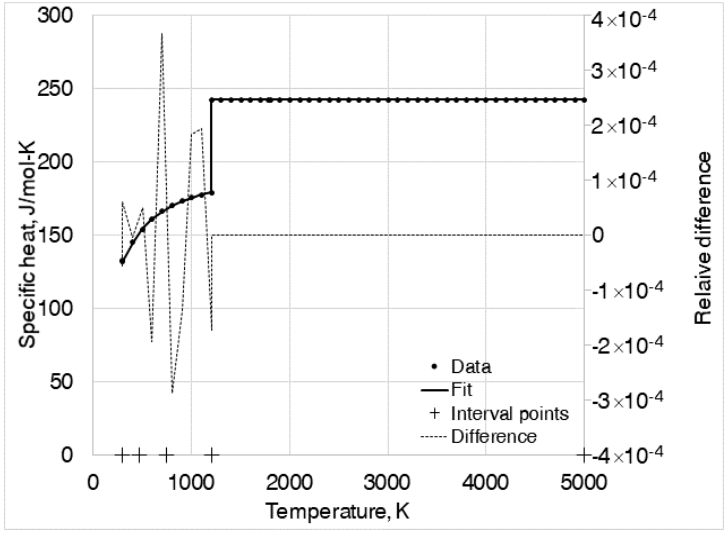


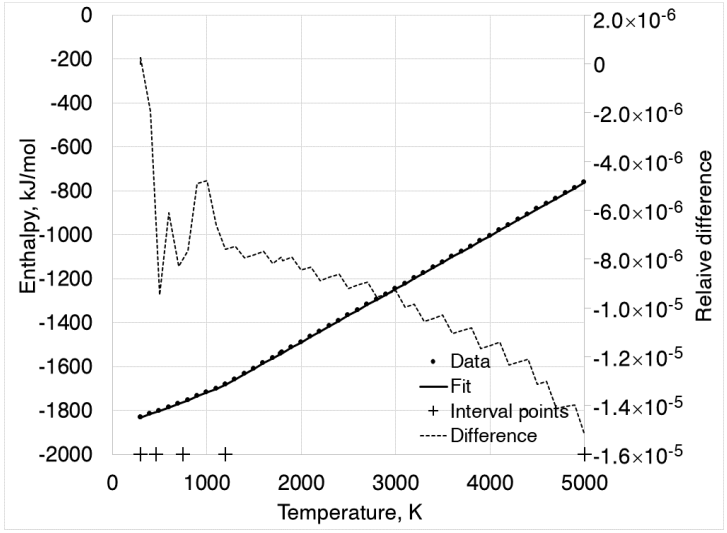


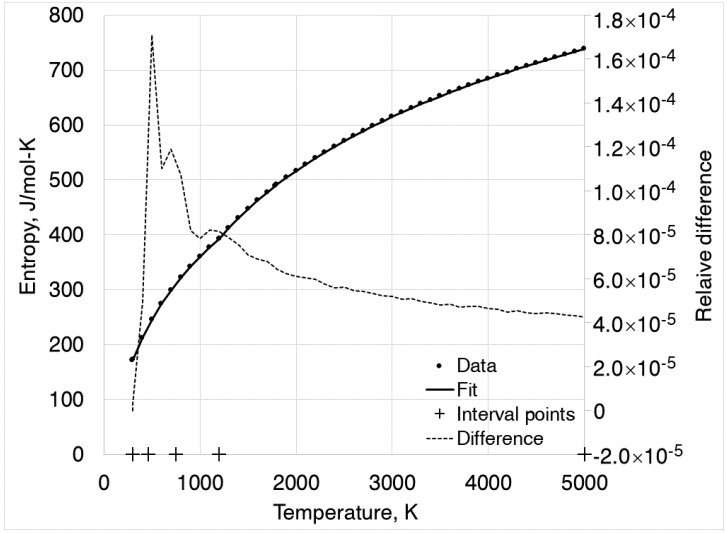


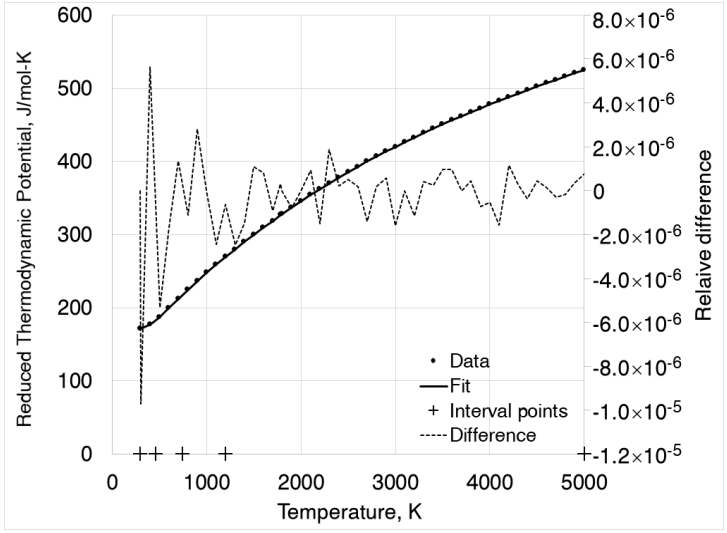
Niobium Oxide (Nb_2O_5) – Nb2O5(l)

MW = 265.8098 g/mol – $\Delta_f H^o$(298.15 K) = -1830.873 kJ/mol – S^o(298.15 K) = 171.488 J/mol-K – p^o = 0.1 MPa



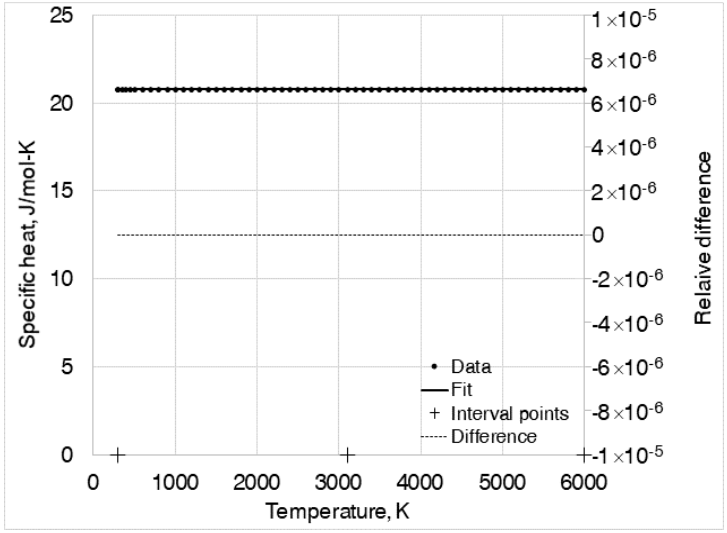


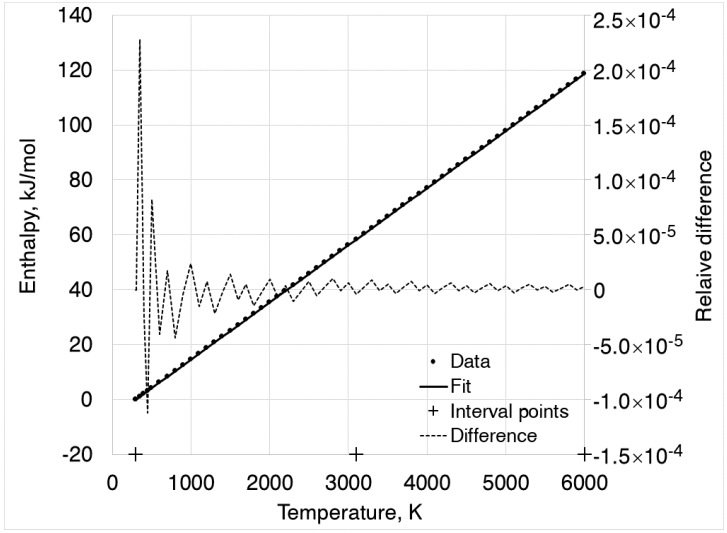


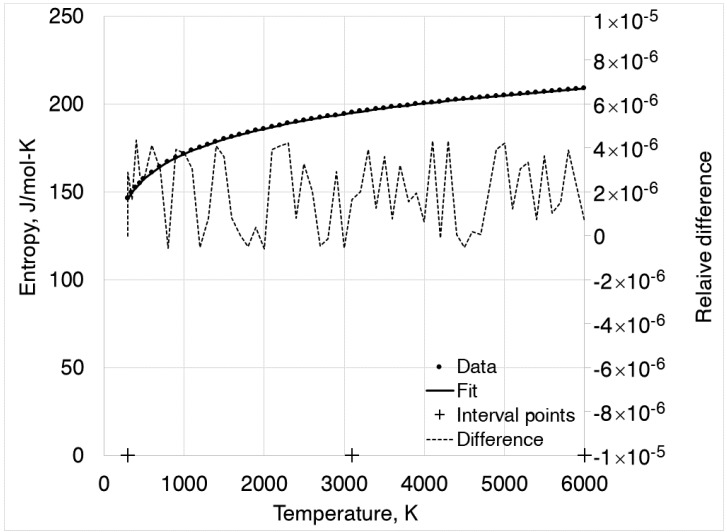


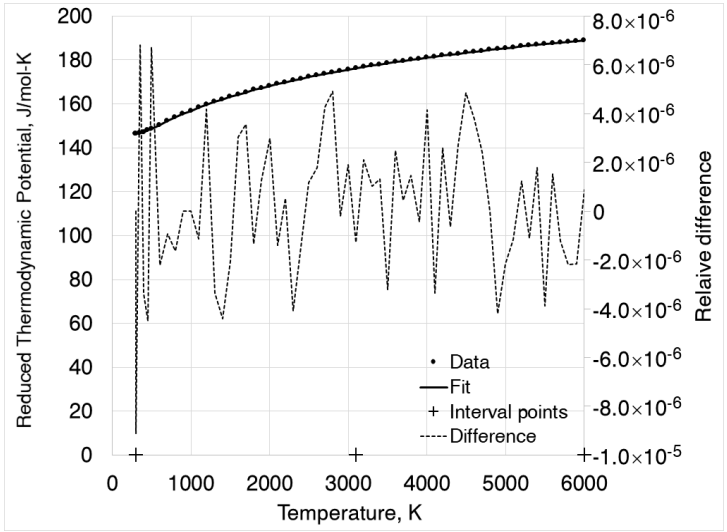
Neon (Ne) – Ne1(ref)

MW = 20.179 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 146.327 J/mol-K – p^o = 0.1 MPa



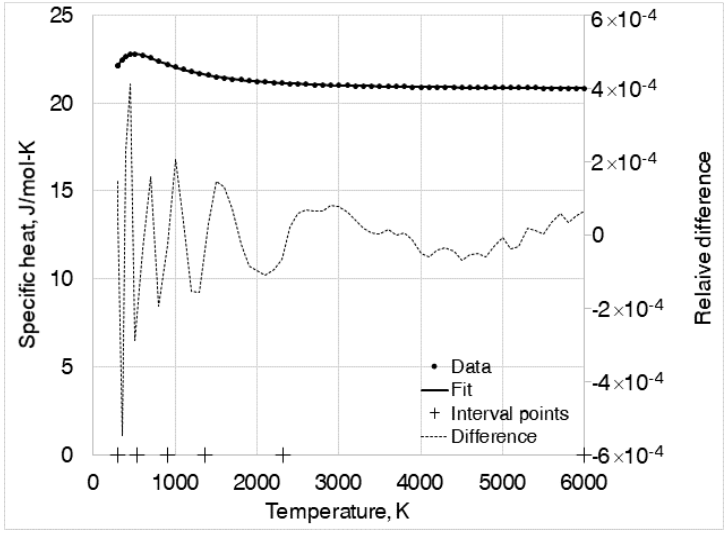


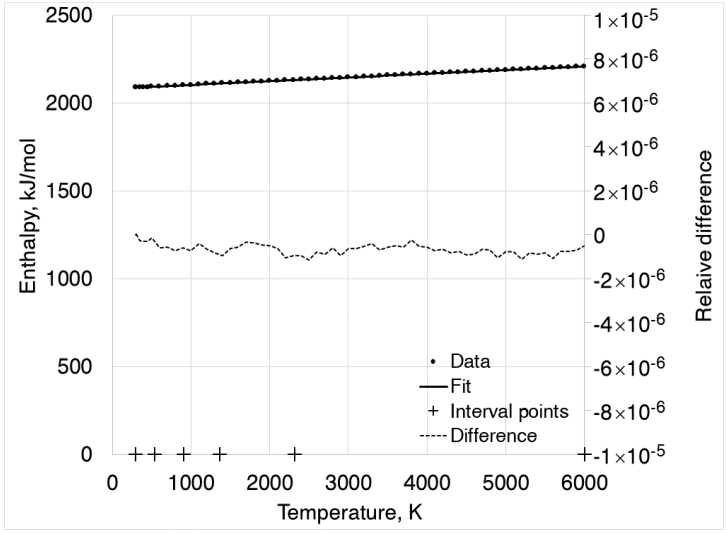


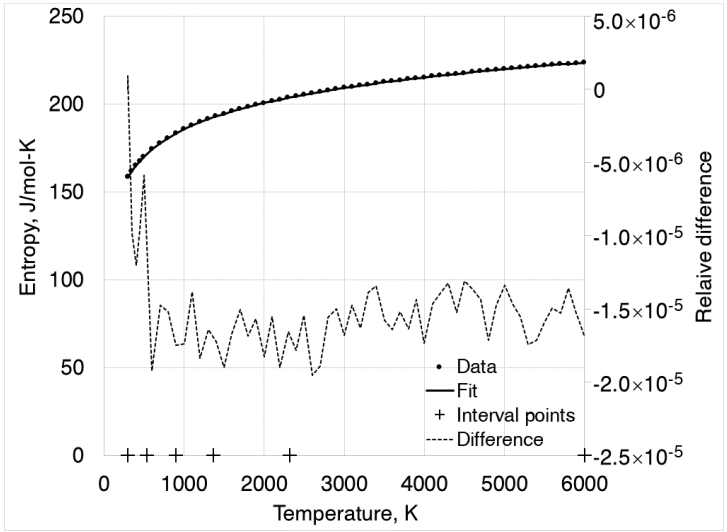


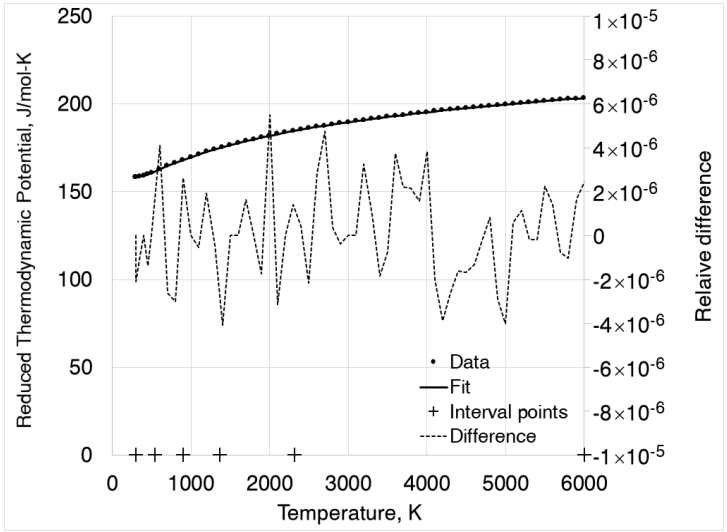
Neon, Ion (Ne^+) – Ne1+(g)

MW = 20.178 g/mol – $\Delta_f H^o$(298.15 K) = 2086.966 kJ/mol – S^o(298.15 K) = 158.307 J/mol-K – p^o = 0.1 MPa



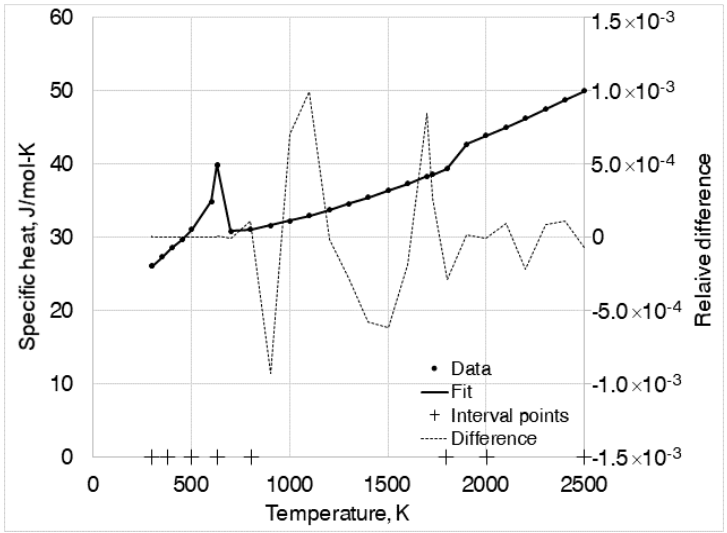


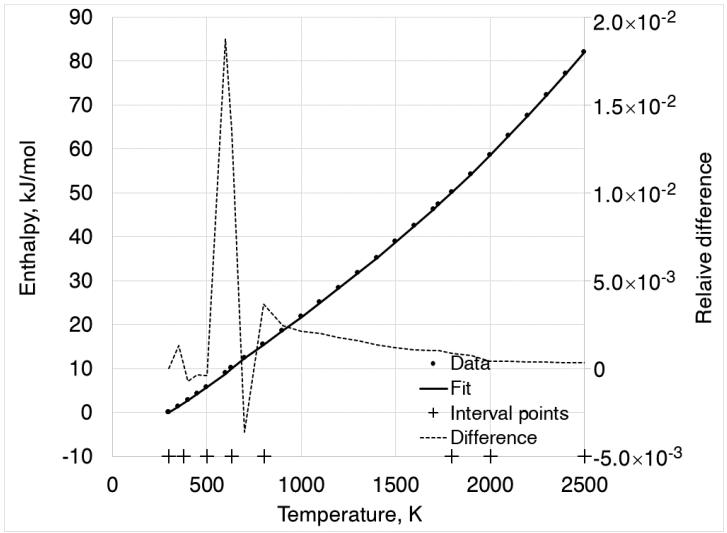


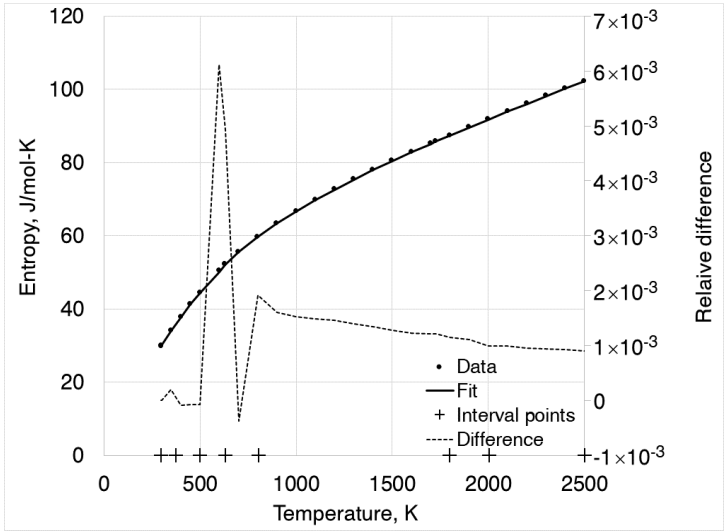


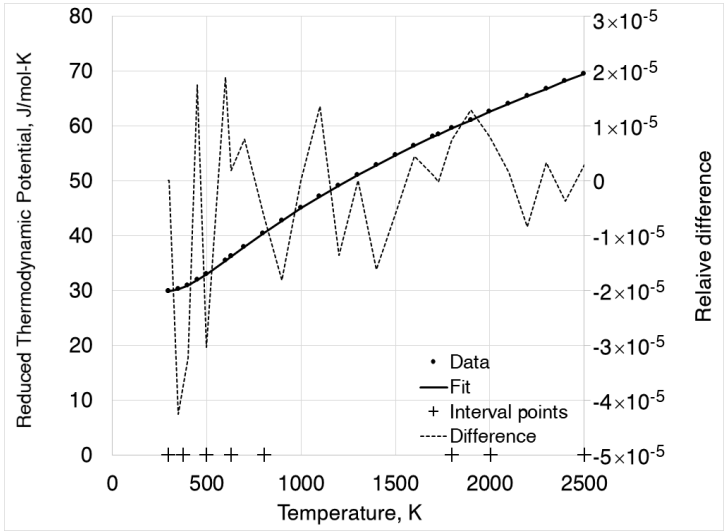
Nickel (Ni) – Ni1(cr)

MW = 58.69 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 29.87 J/mol-K – p^o = 0.1 MPa



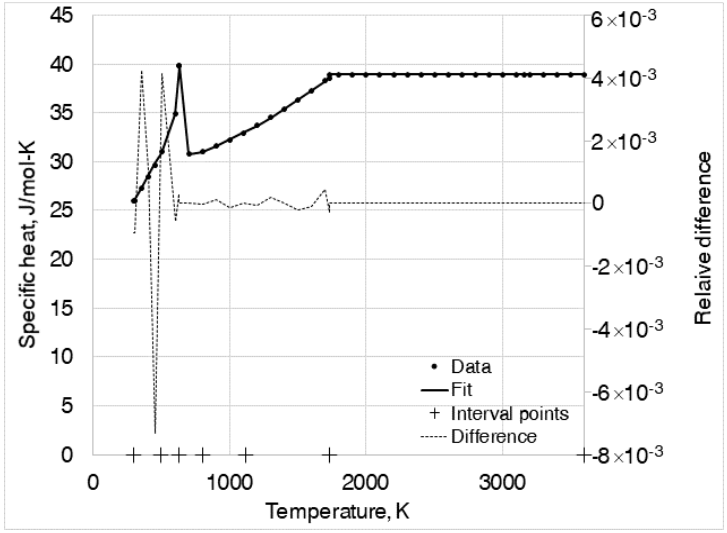


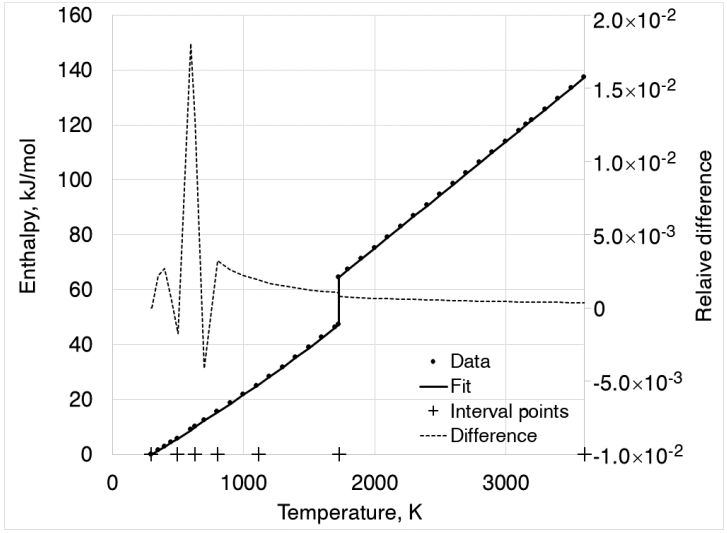


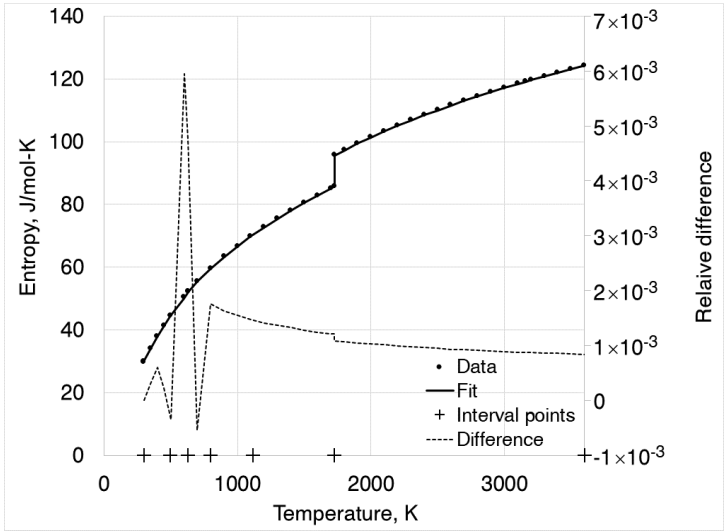


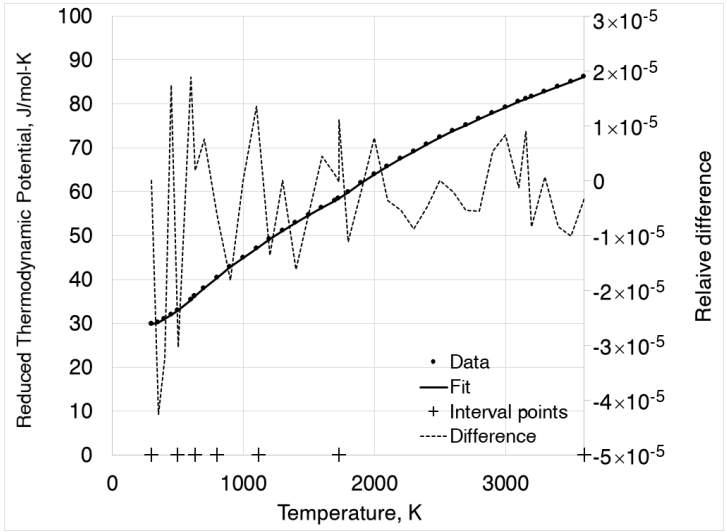
Nickel (Ni) – Ni1(cr,l)

MW = 58.69 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 29.87 J/mol-K – p^o = 0.1 MPa



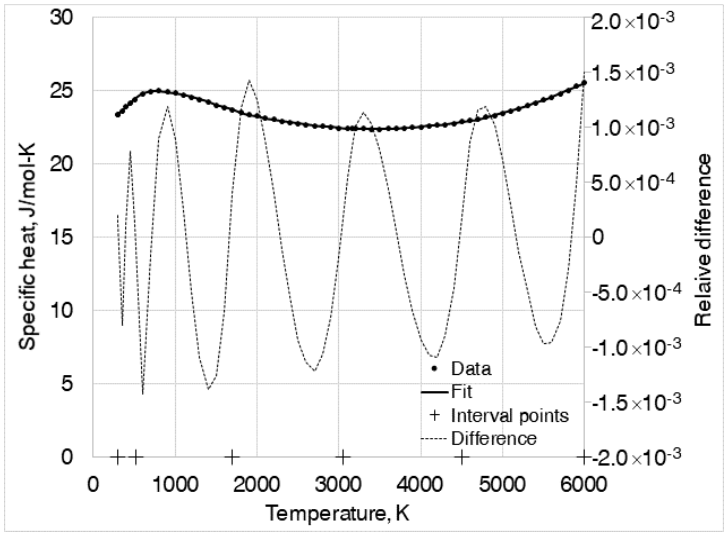


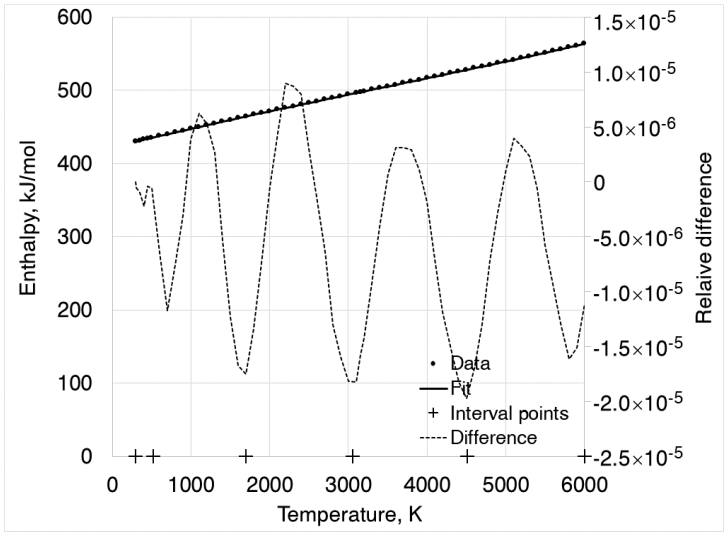


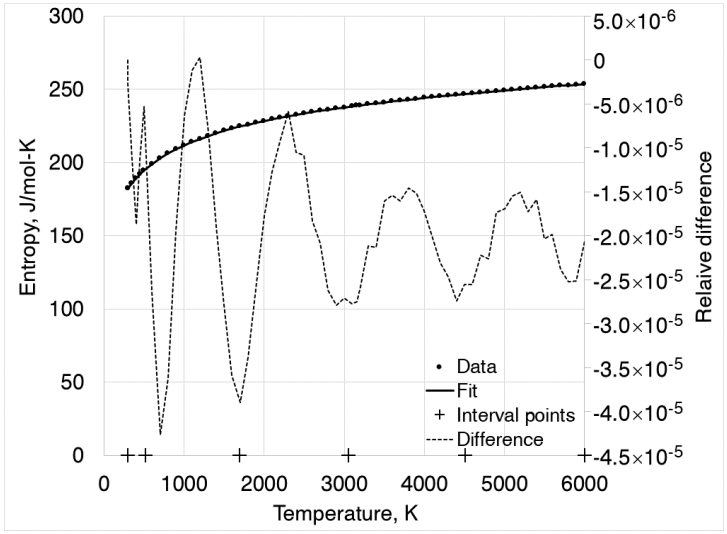


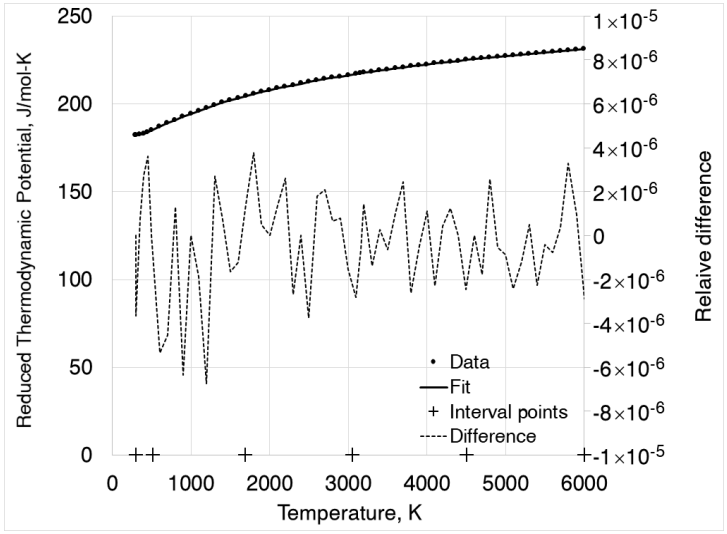
Nickel (Ni) – Ni1(g)

MW = 58.69 g/mol – $\Delta_f H^o$(298.15 K) = 430.115 kJ/mol – S^o(298.15 K) = 182.19 J/mol-K – p^o = 0.1 MPa



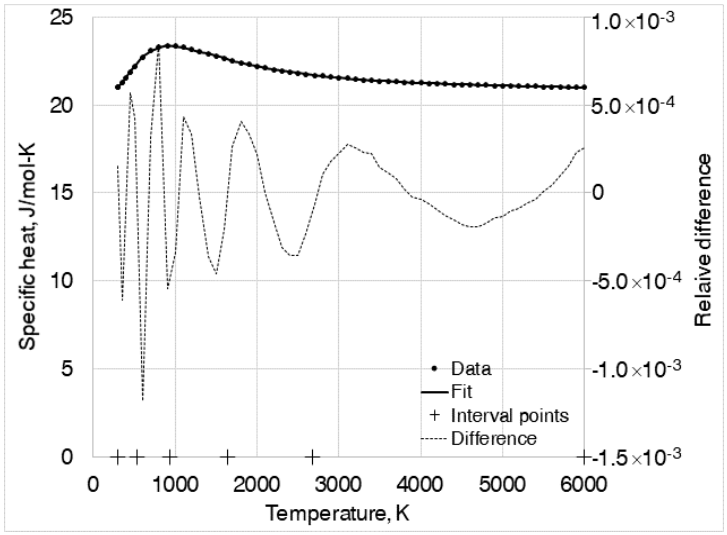


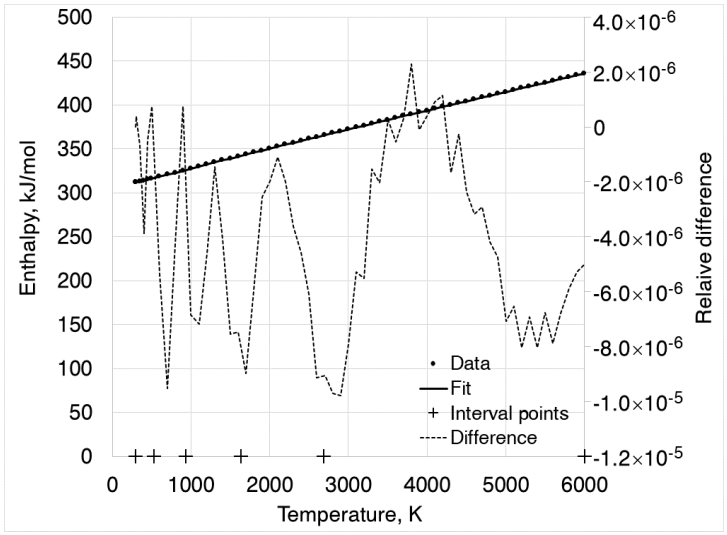


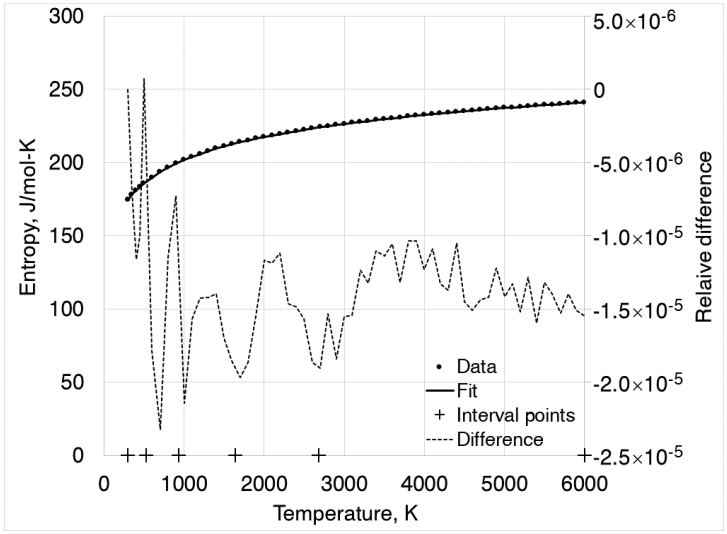


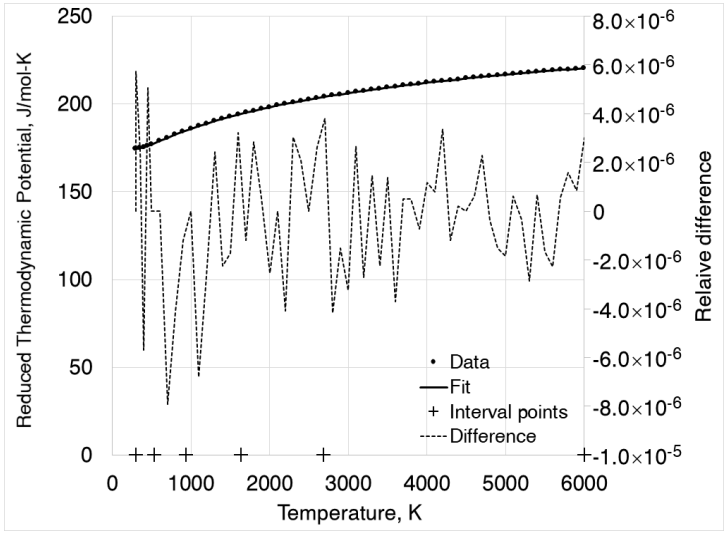
Nickel, Ion (Ni^-) – Ni1-(g)

MW = 58.69 g/mol – $\Delta_f H^o$(298.15 K) = 311.764 kJ/mol – S^o(298.15 K) = 174.577 J/mol-K – p^o = 0.1 MPa



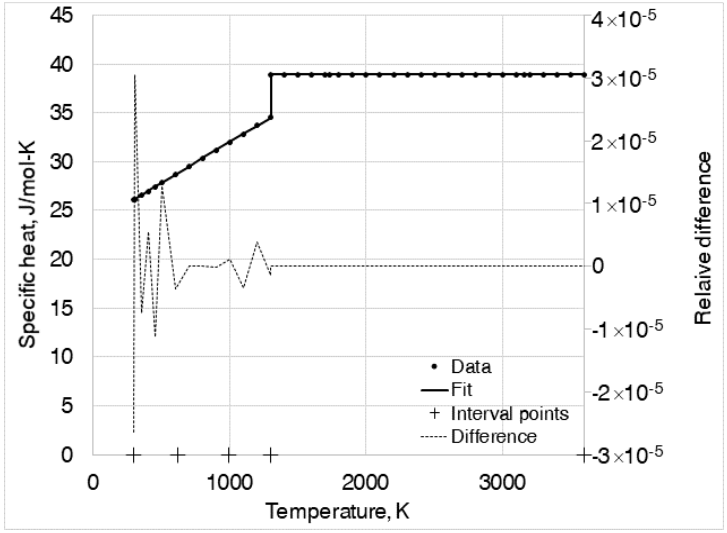


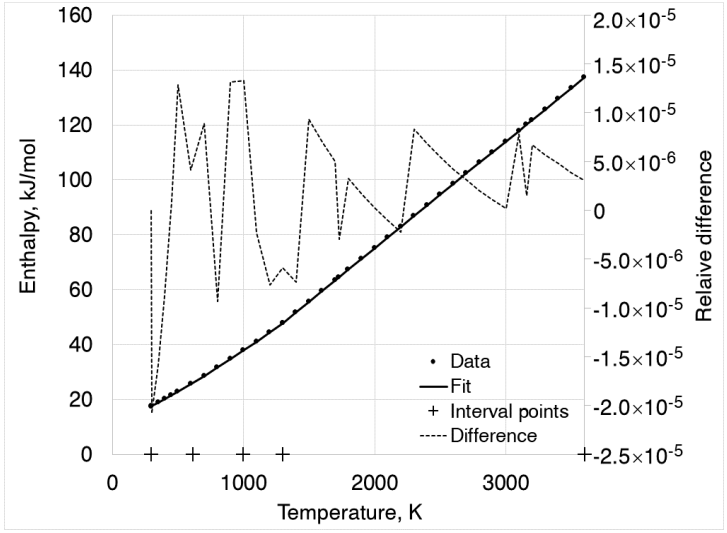


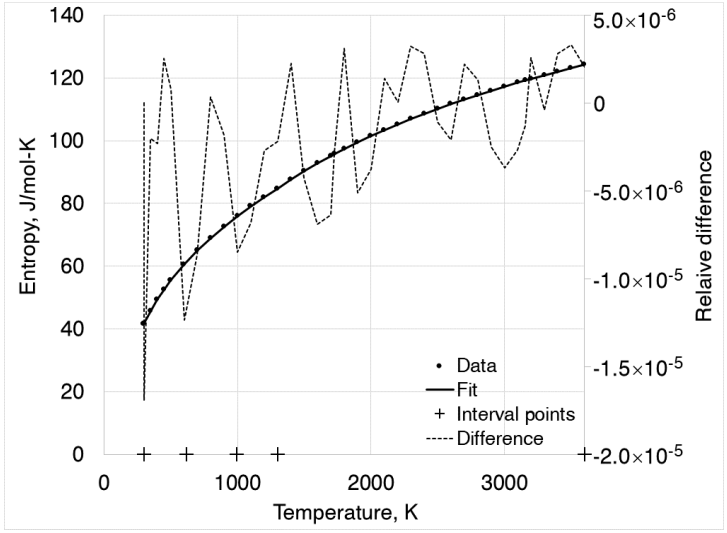


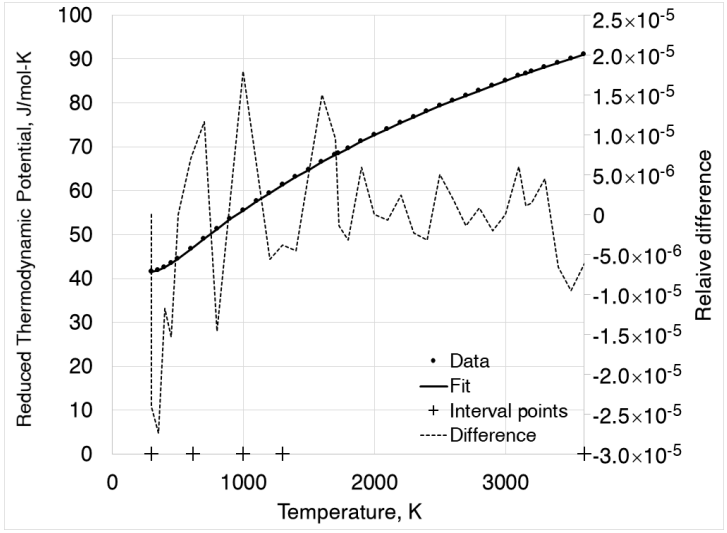
Nickel (Ni) – Ni1(l)

MW = 58.69 g/mol – $\Delta_f H^o$(298.15 K) = 17.479 kJ/mol – S^o(298.15 K) = 41.487 J/mol-K – p^o = 0.1 MPa



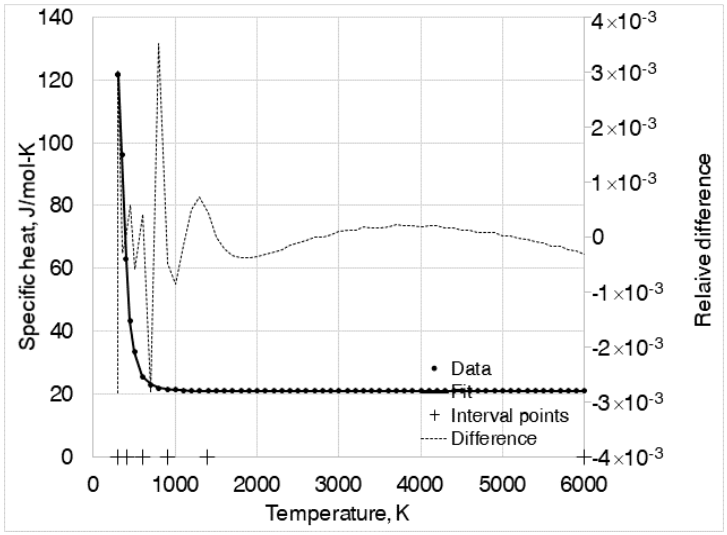


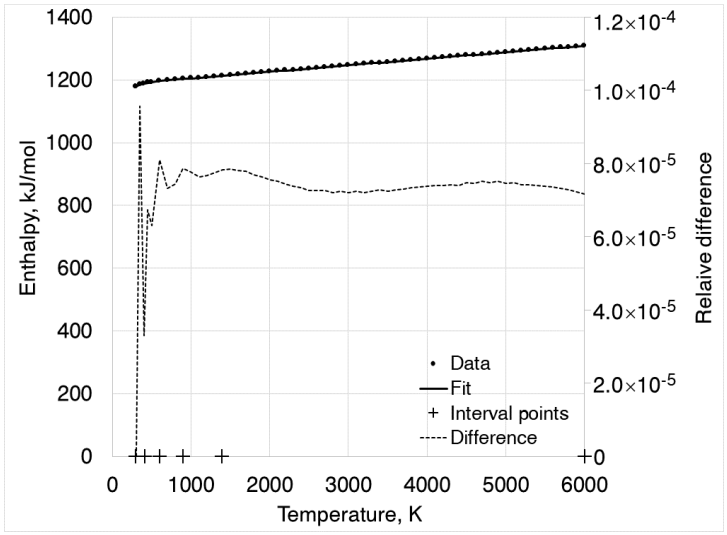


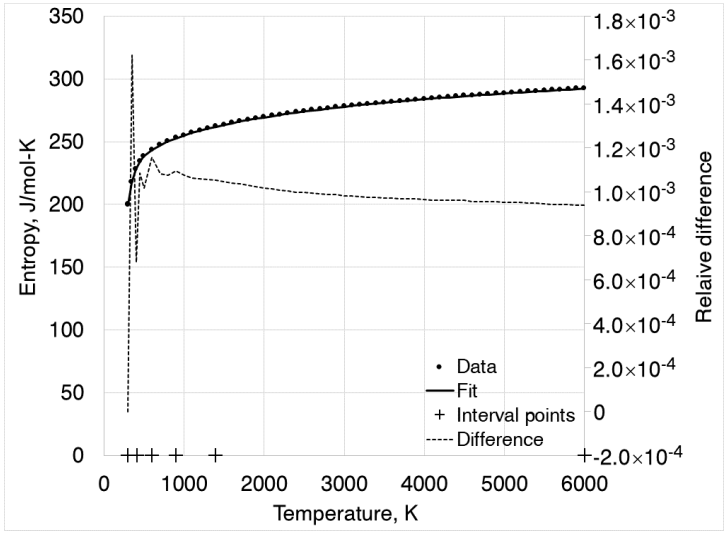


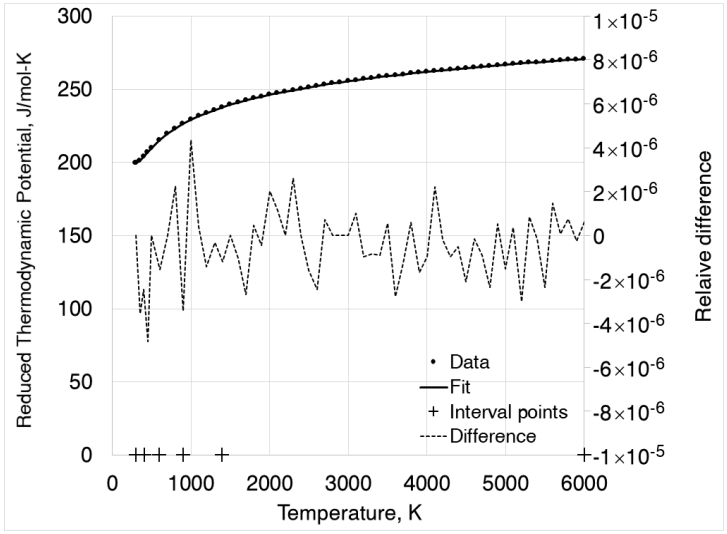
Nickel, Ion (Ni^+) – Ni1+(g)

MW = 58.69 g/mol – $\Delta_f H^o$(298.15 K) = 1179.003 kJ/mol – S^o(298.15 K) = 199.724 J/mol-K – p^o = 0.1 MPa



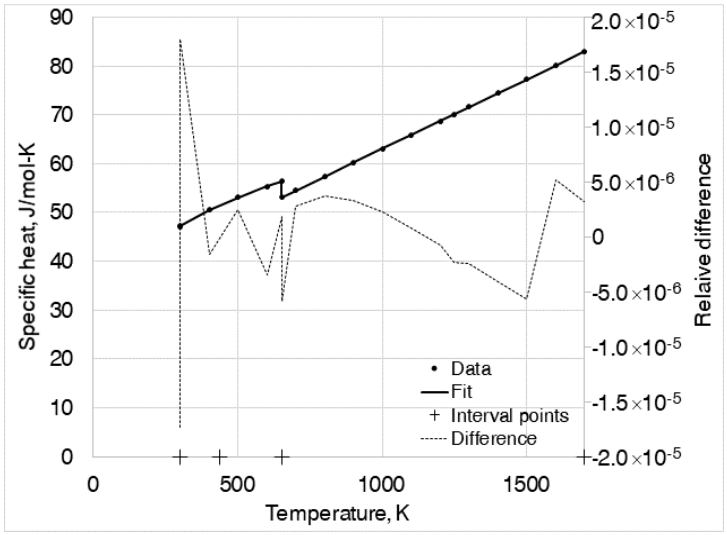


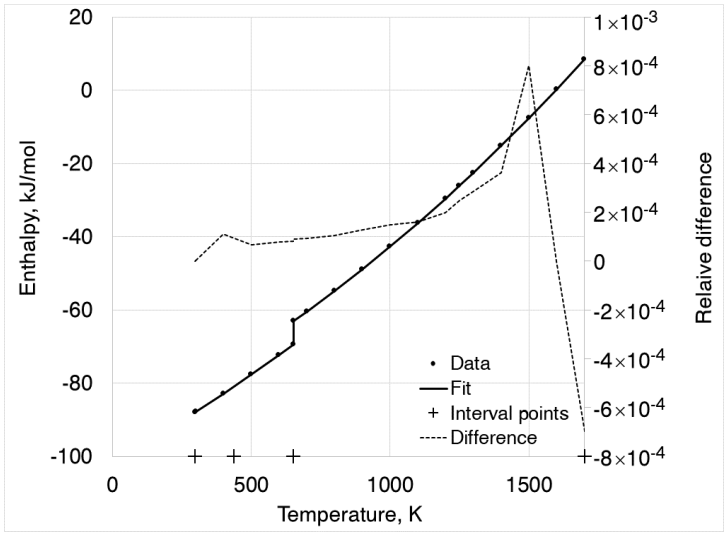


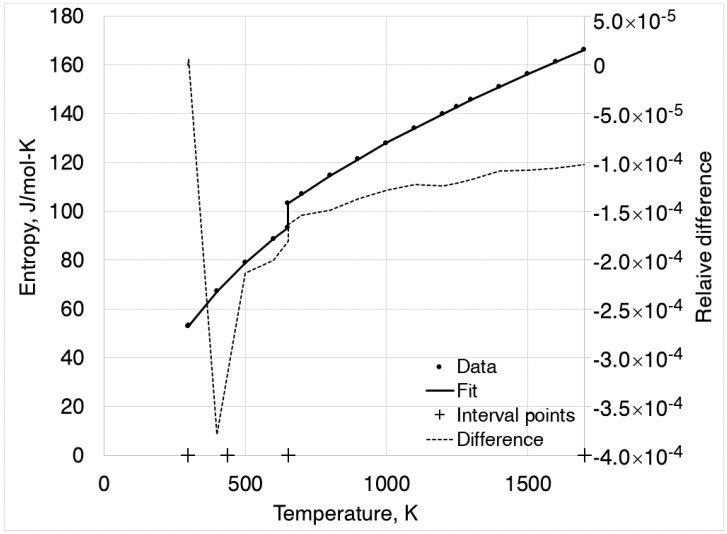


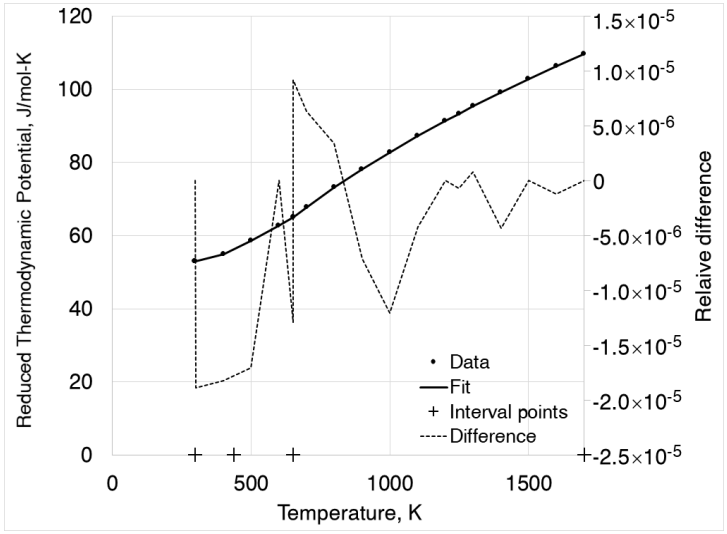
Nickel Sulfide (NiS) – Ni1S1(cr)

MW = 90.75 g/mol – $\Delta_f H^o$(298.15 K) = -87.864 kJ/mol – S^o(298.15 K) = 52.992 J/mol-K – p^o = 0.1 MPa



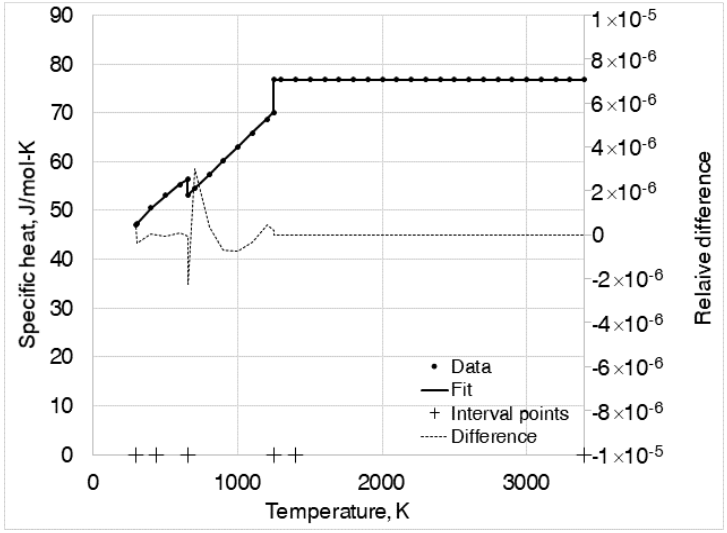


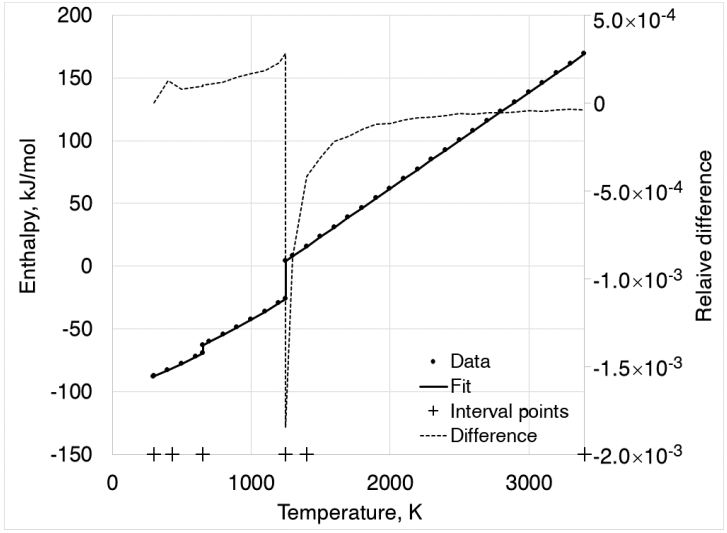


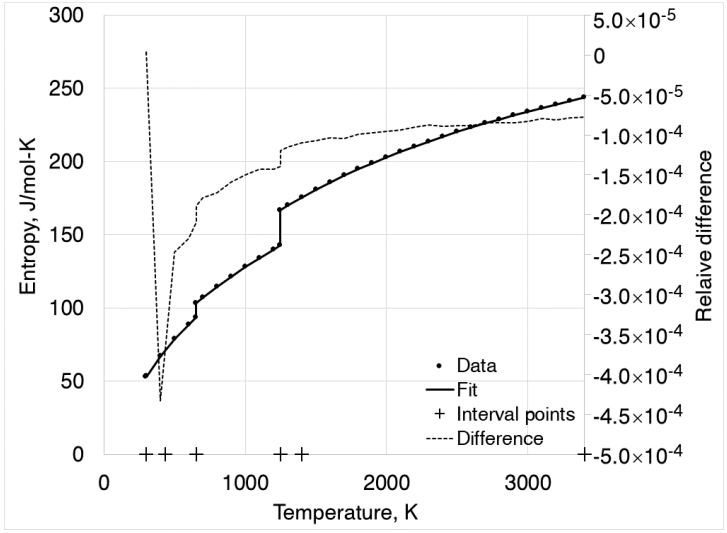


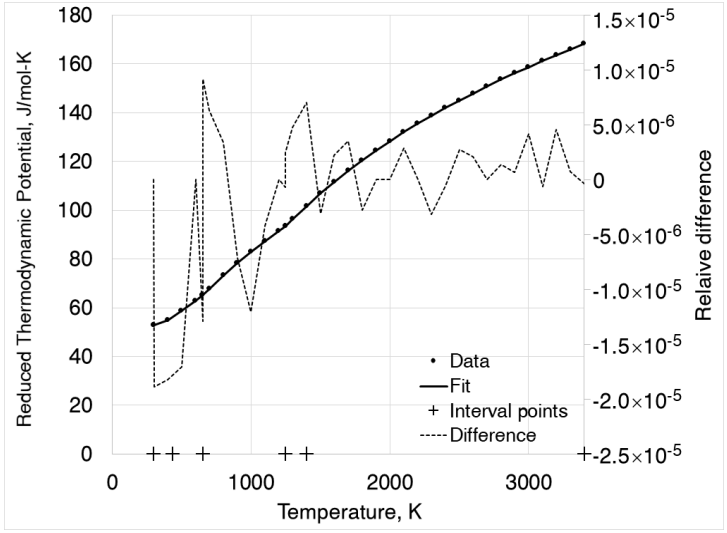
Nickel Sulfide (NiS) – Ni1S1(cr,l)

MW = 90.75 g/mol – $\Delta_f H^o$(298.15 K) = -87.864 kJ/mol – S^o(298.15 K) = 52.992 J/mol-K – p^o = 0.1 MPa



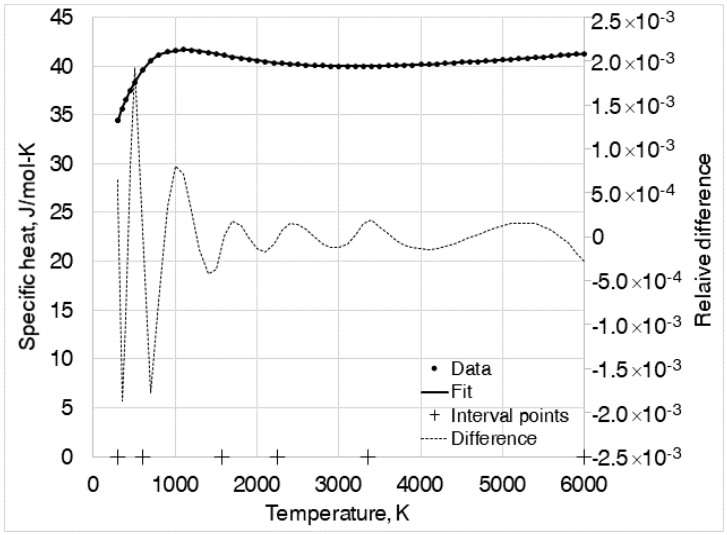


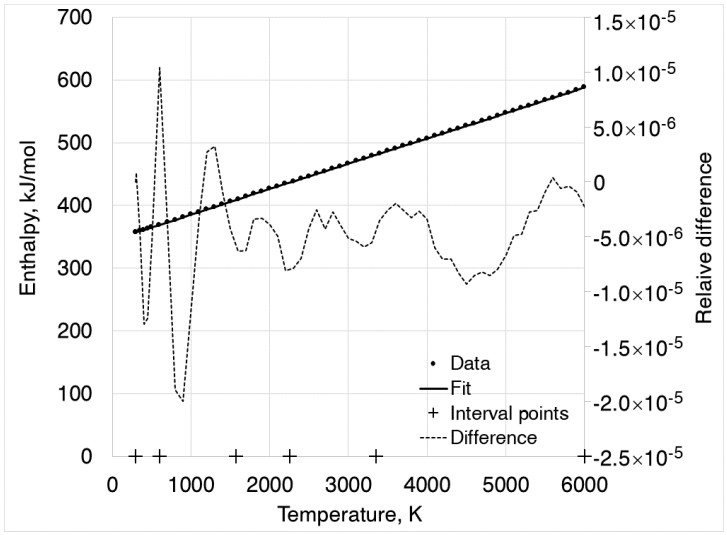


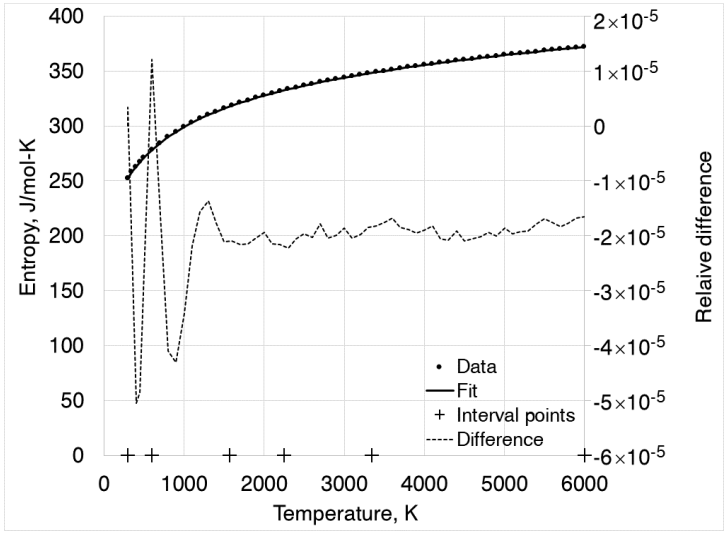


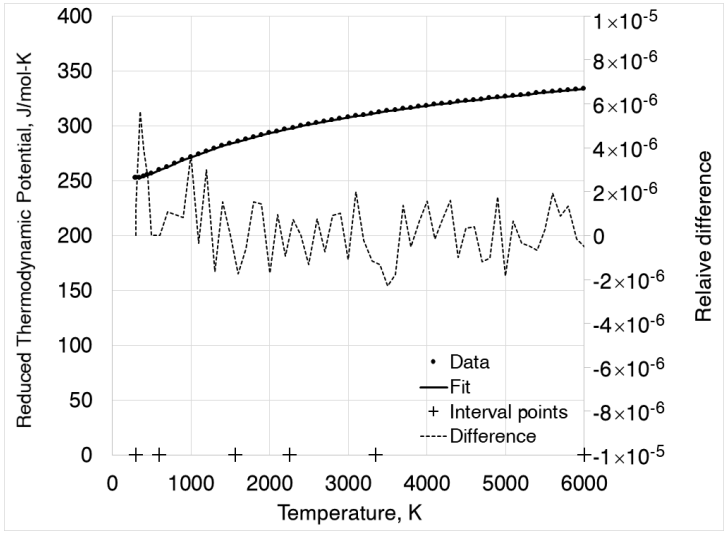
Nickel Sulfide (NiS) – Ni1S1(g)

MW = 90.75 g/mol – $\Delta_f H^o$(298.15 K) = 357.419 kJ/mol – S^o(298.15 K) = 252.282 J/mol-K – p^o = 0.1 MPa



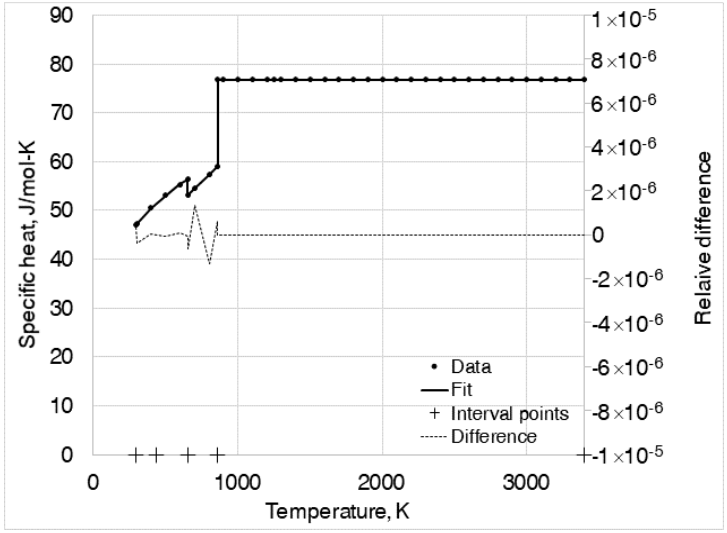


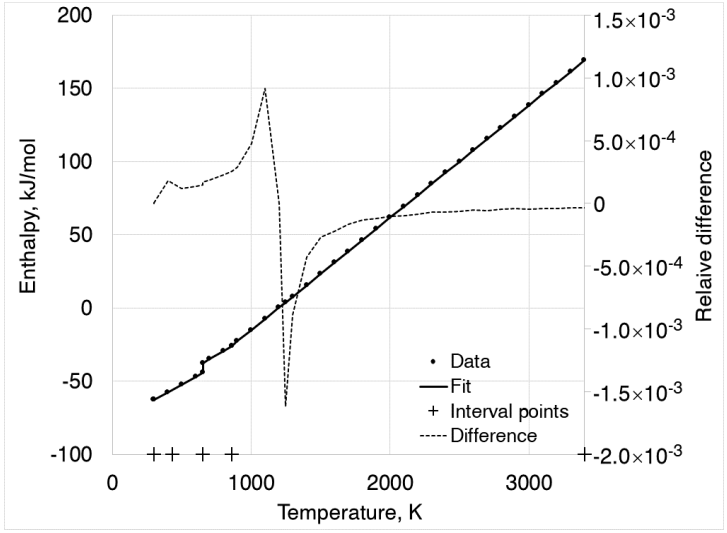


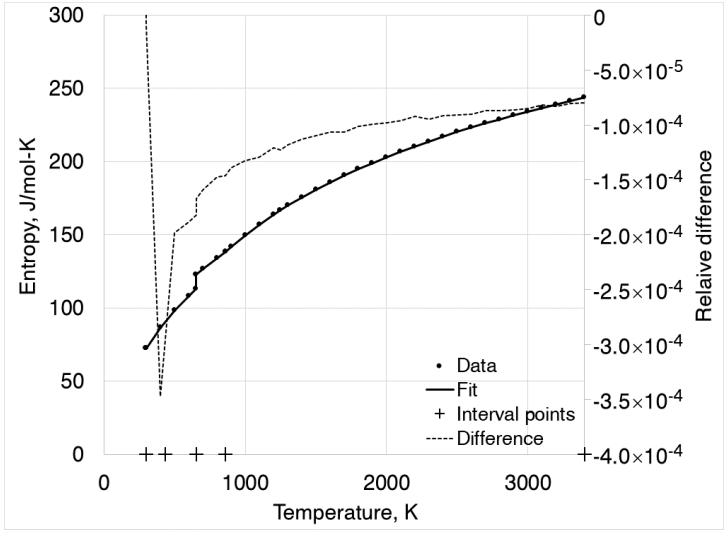


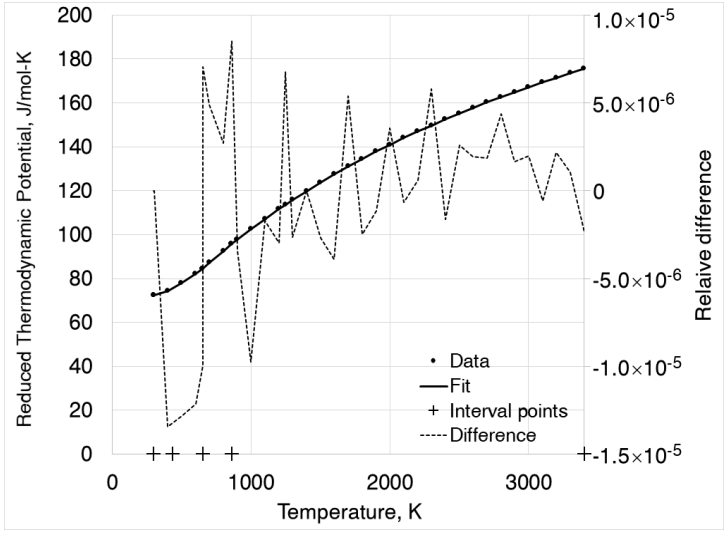
Nickel Sulfide (NiS) – Ni1S1(l)

MW = 90.75 g/mol – $\Delta_f H^o$(298.15 K) = -62.474 kJ/mol – S^o(298.15 K) = 72.441 J/mol-K – p^o = 0.1 MPa



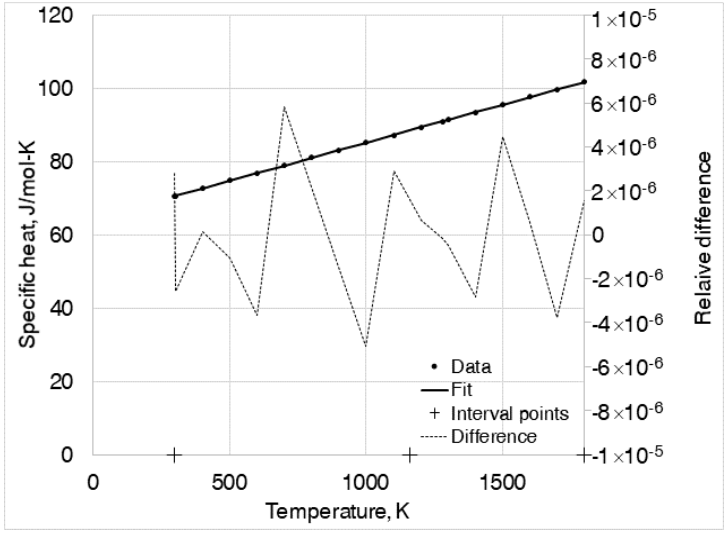


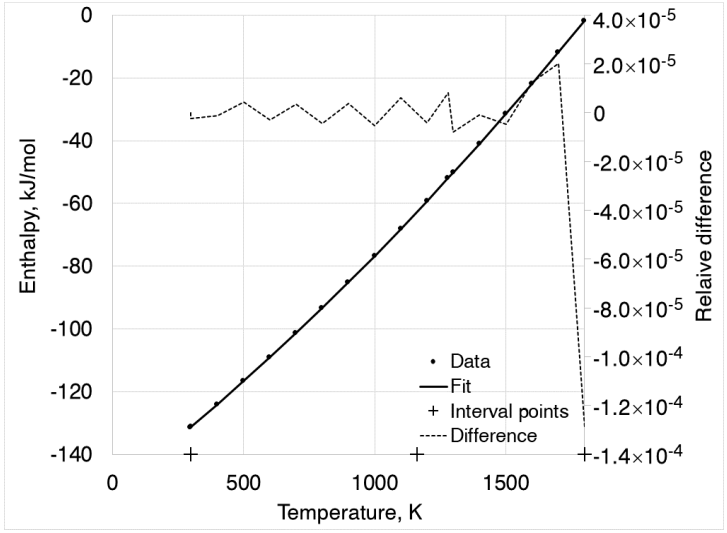


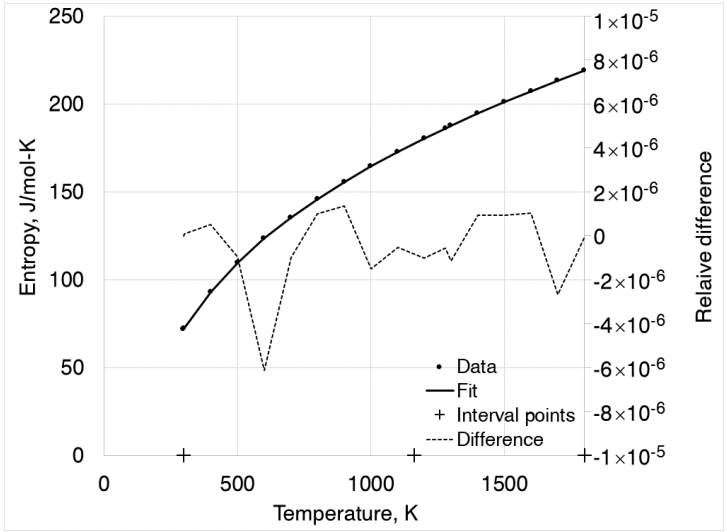


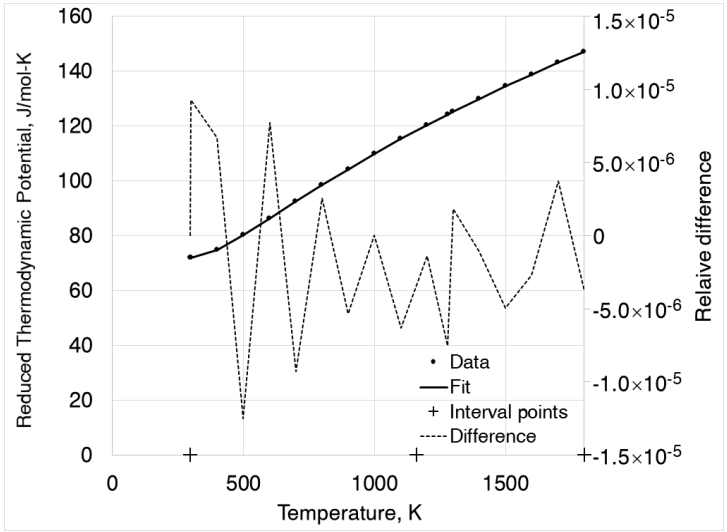
Nickel Sulfide (NiS_2) – Ni1S2(cr)

MW = 122.81 g/mol – $\Delta_f H^o$(298.15 K) = -131.378 kJ/mol – S^o(298.15 K) = 71.965 J/mol-K – p^o = 0.1 MPa



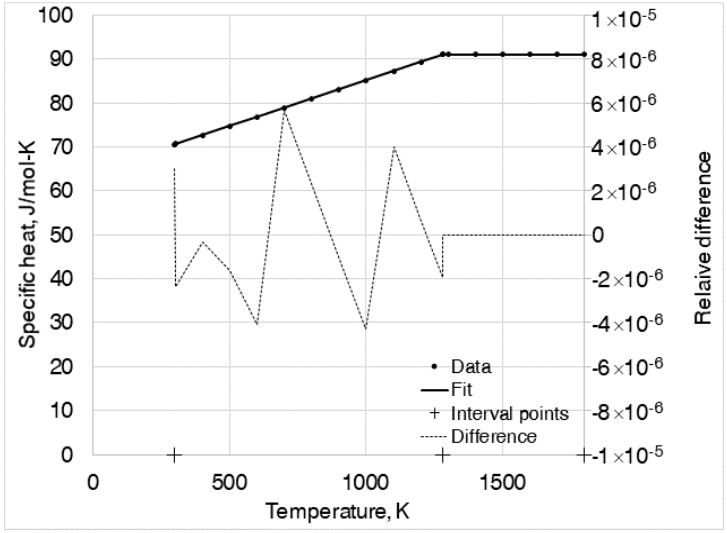


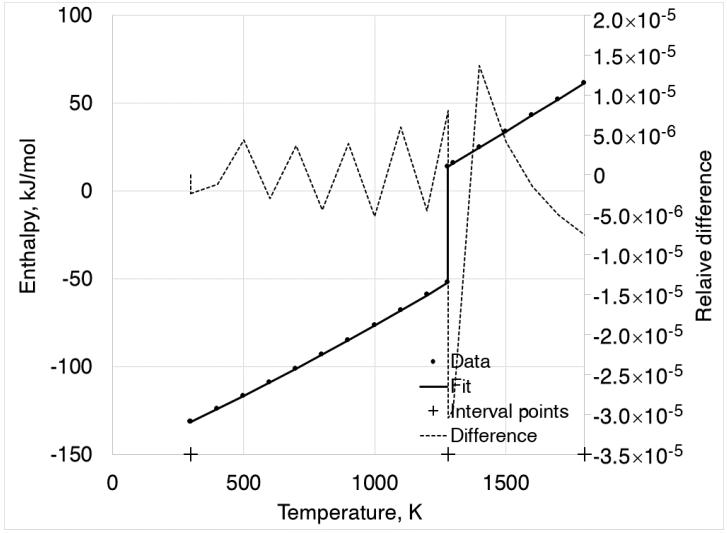


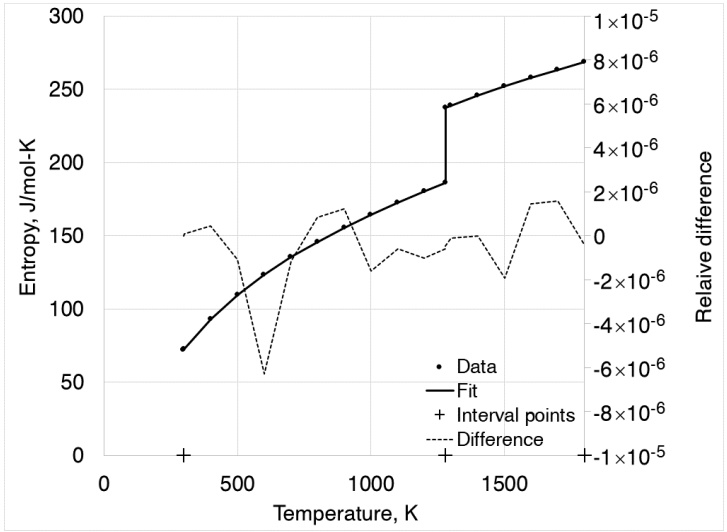


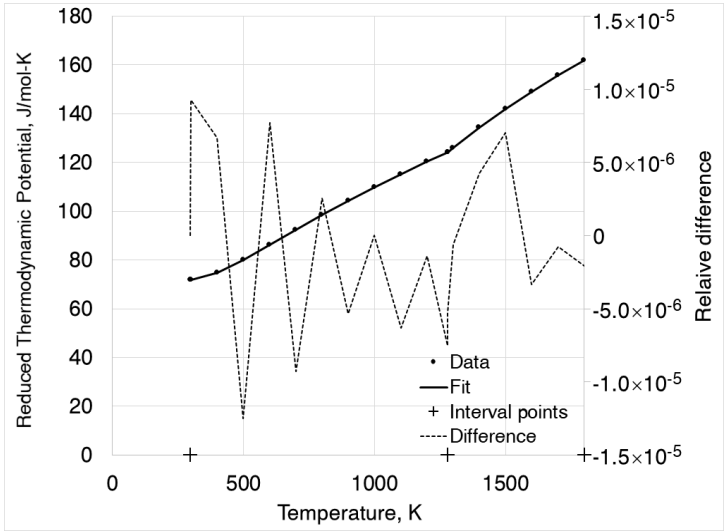
Nickel Sulfide (NiS_2) – Ni1S2(cr,l)

MW = 122.81 g/mol – $\Delta_f H^o$(298.15 K) = -131.378 kJ/mol – S^o(298.15 K) = 71.965 J/mol-K – p^o = 0.1 MPa



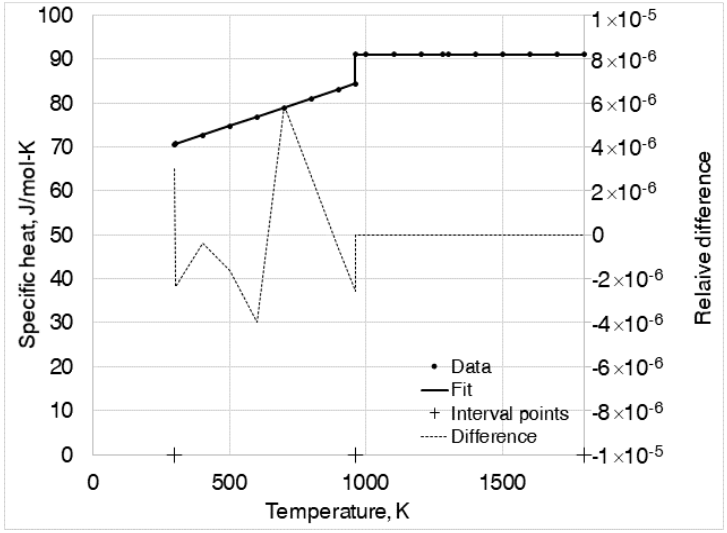


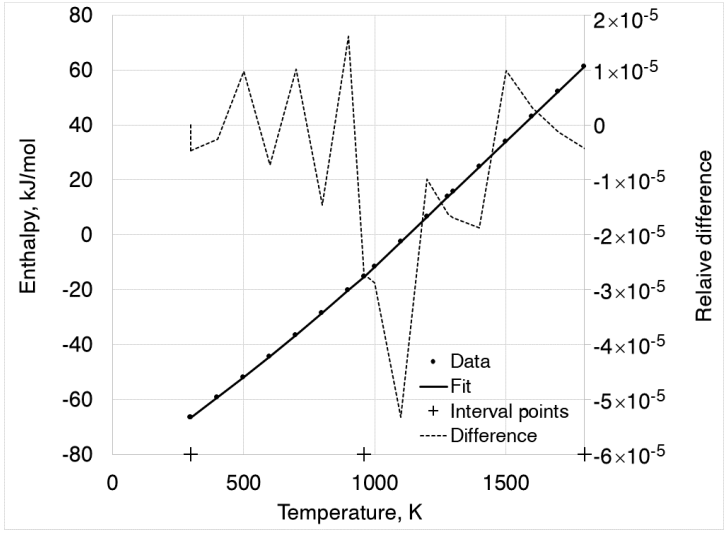


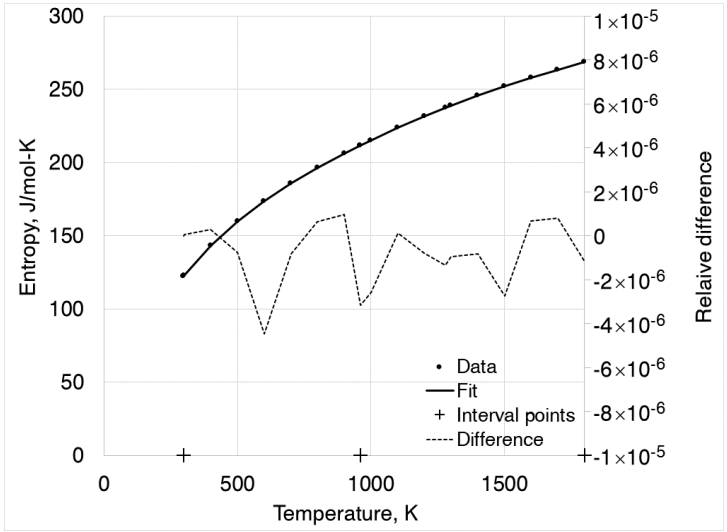


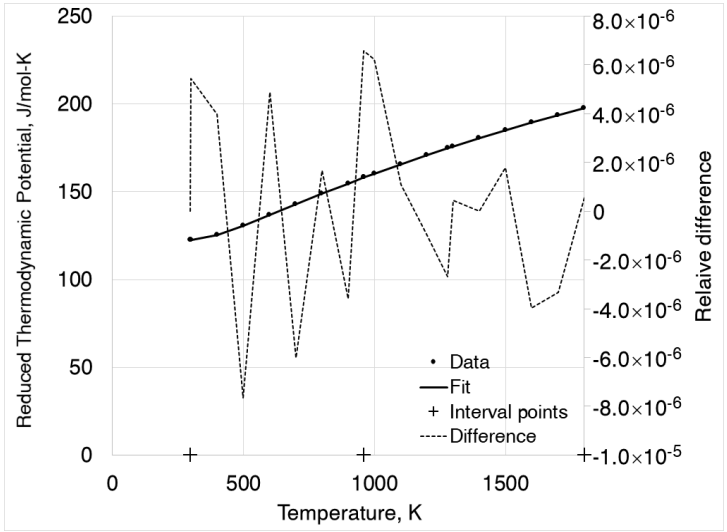
Nickel Sulfide (NiS_2) – Ni1S2(l)

MW = 122.81 g/mol – $\Delta_f H^o$(298.15 K) = -66.568 kJ/mol – S^o(298.15 K) = 122.427 J/mol-K – p^o = 0.1 MPa



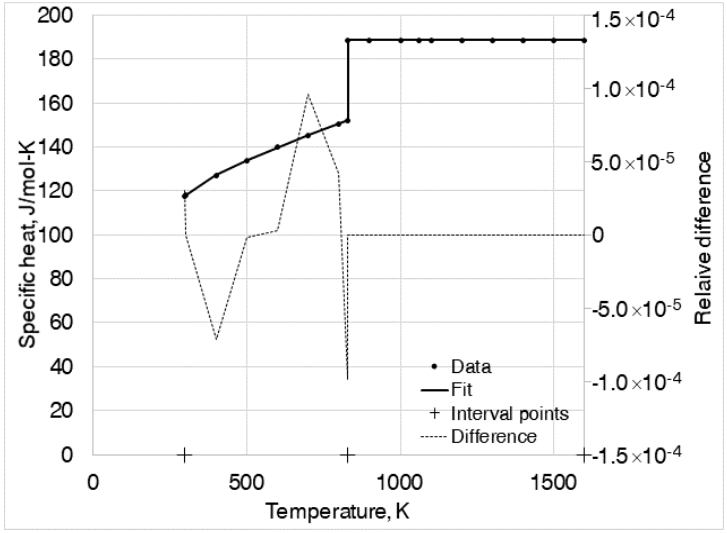


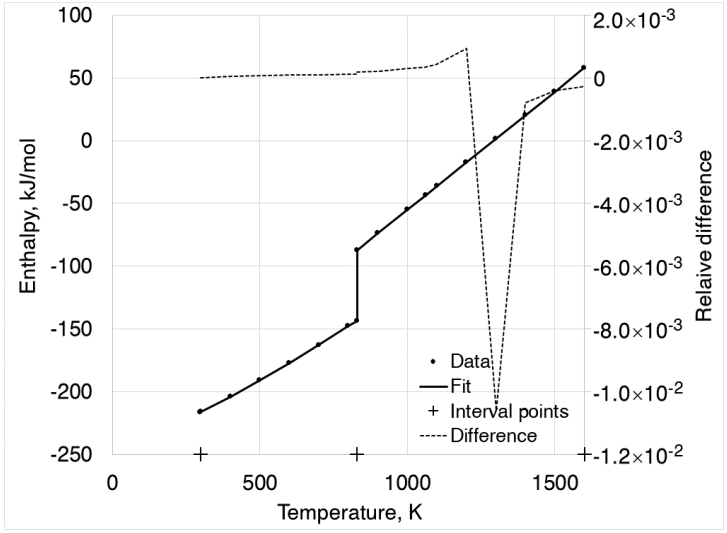


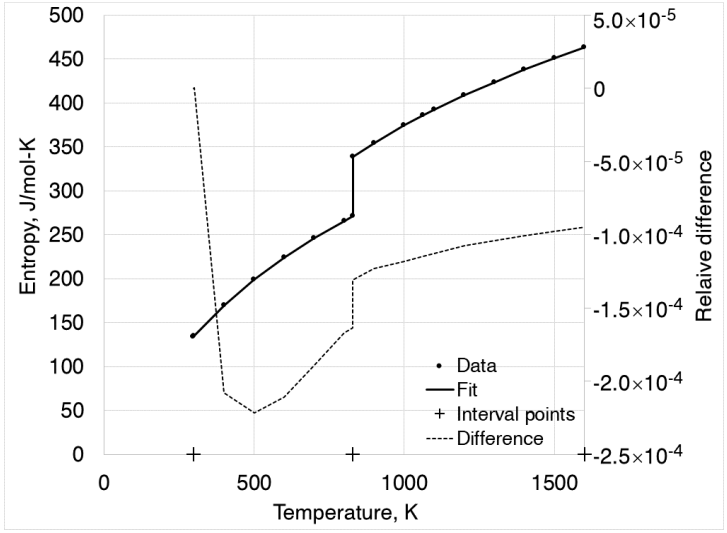


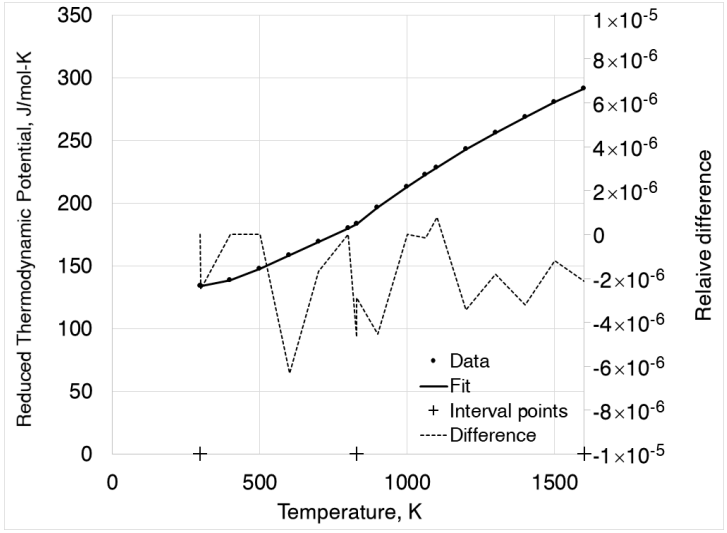
Nickel Sulfide (Ni_3S_2) – Ni3S2(cr)

MW = 240.19 g/mol – $\Delta_f H^o$(298.15 K) = -216.313 kJ/mol – S^o(298.15 K) = 133.887 J/mol-K – p^o = 0.1 MPa



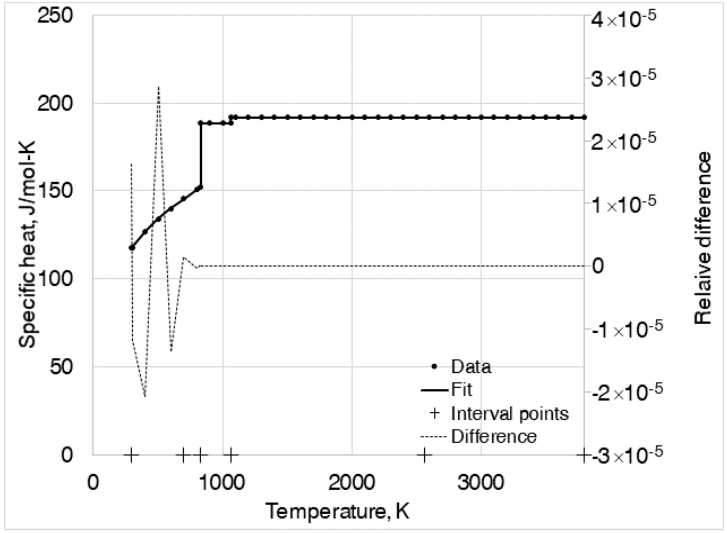


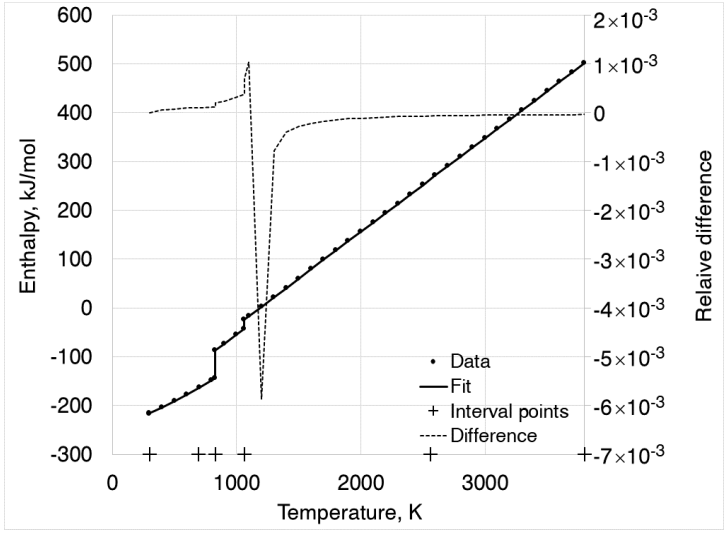


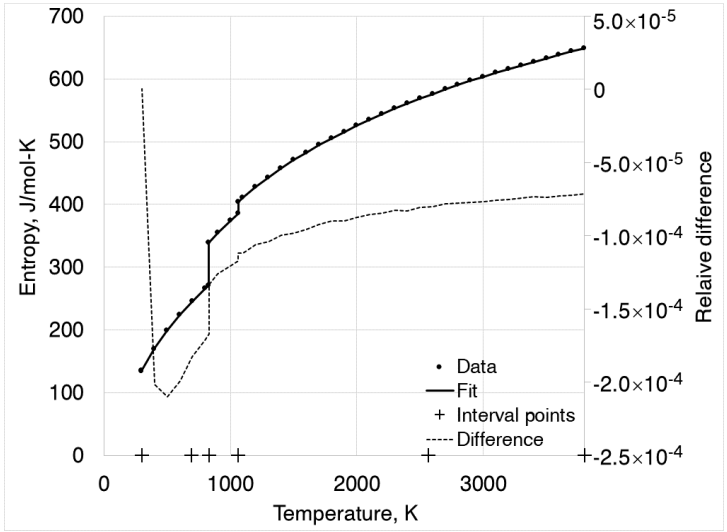


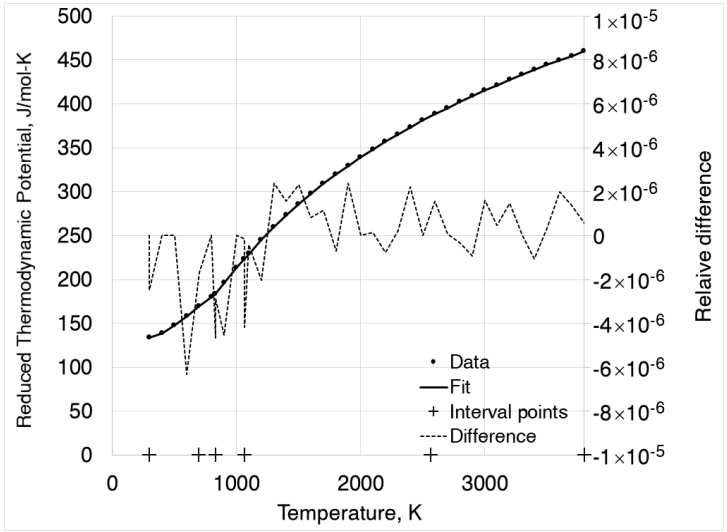
Nickel Sulfide (Ni_3S_2) – Ni3S2(cr,l)

MW = 240.19 g/mol – $\Delta_f H^o$(298.15 K) = -216.313 kJ/mol – S^o(298.15 K) = 133.887 J/mol-K – p^o = 0.1 MPa



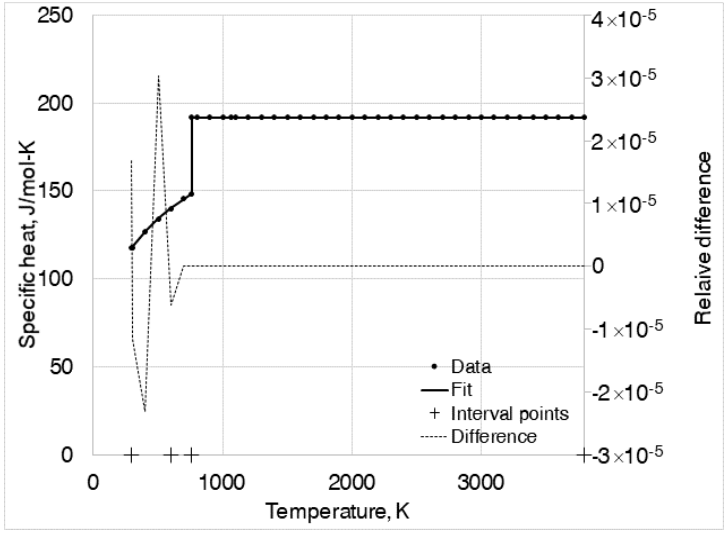


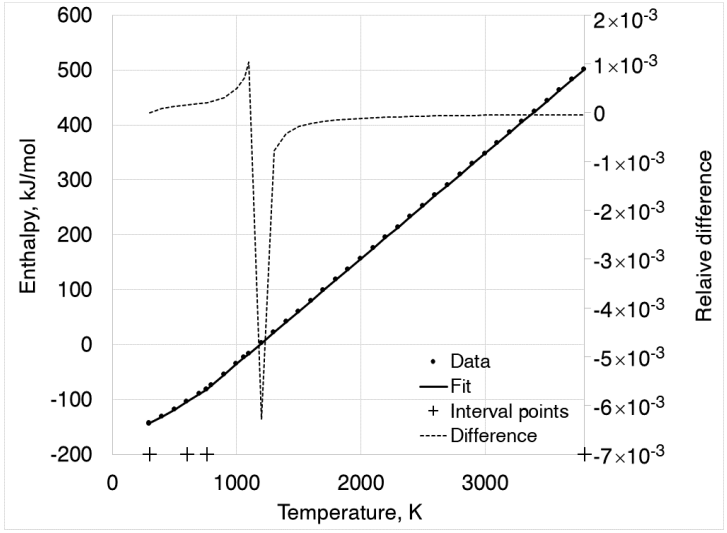


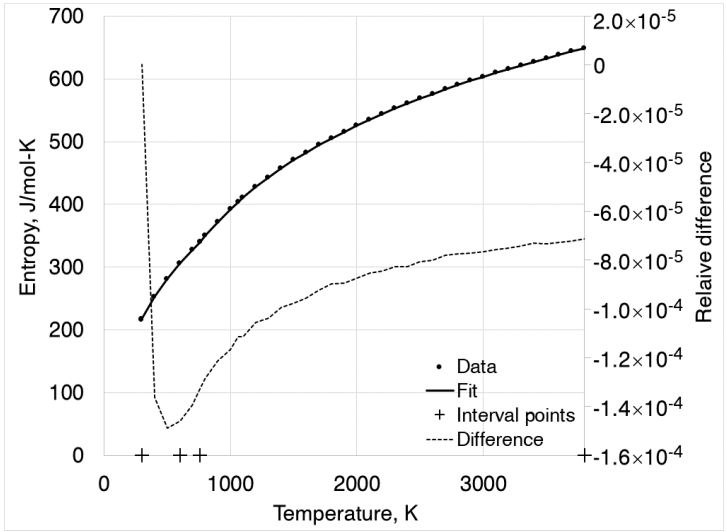


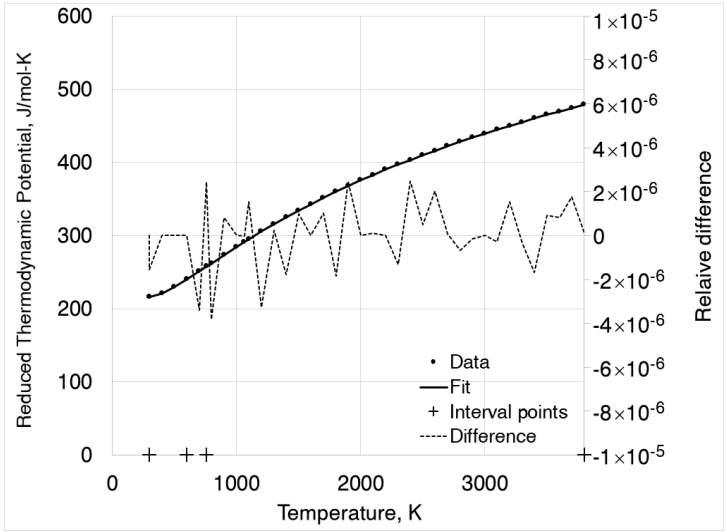
Nickel Sulfide (Ni_3S_2) – Ni3S2(l)

MW = 240.19 g/mol – $\Delta_f H^o$(298.15 K) = -143.926 kJ/mol – S^o(298.15 K) = 215.93 J/mol-K – p^o = 0.1 MPa



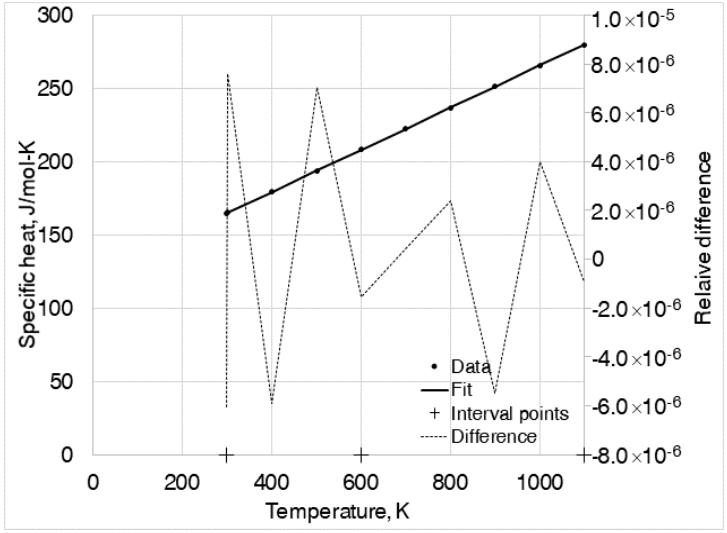


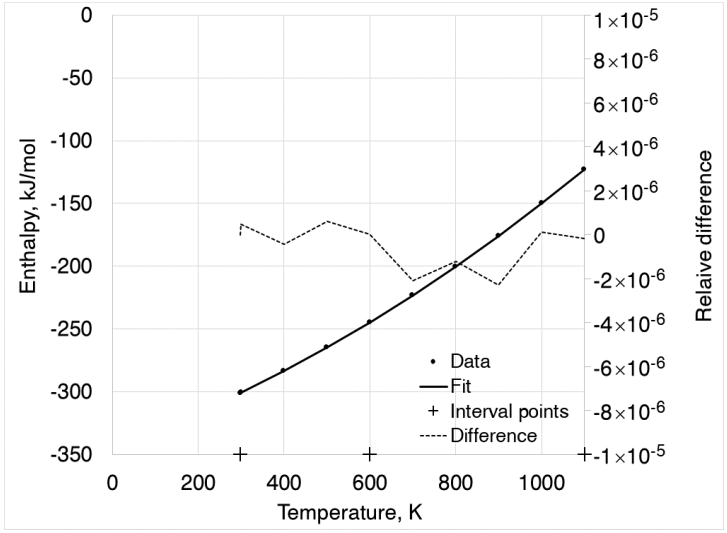


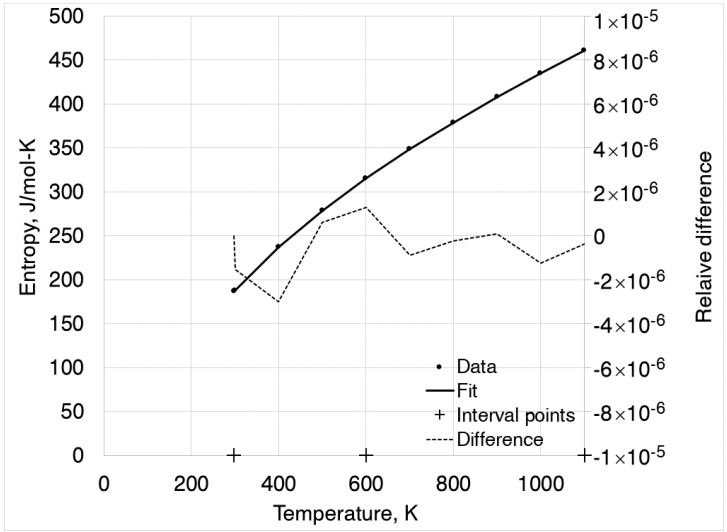


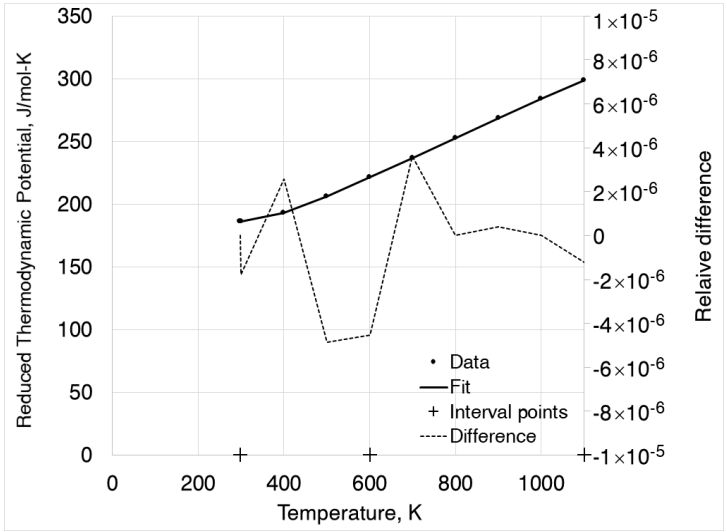
Nickel Sulfide (Ni_3S_4) – Ni3S4(cr)

MW = 304.31 g/mol – $\Delta_f H^o$(298.15 K) = -301.115 kJ/mol – S^o(298.15 K) = 186.481 J/mol-K – p^o = 0.1 MPa



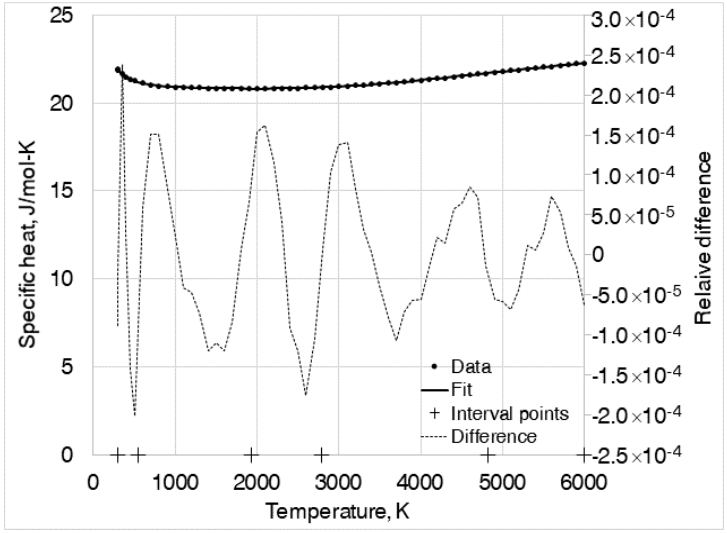


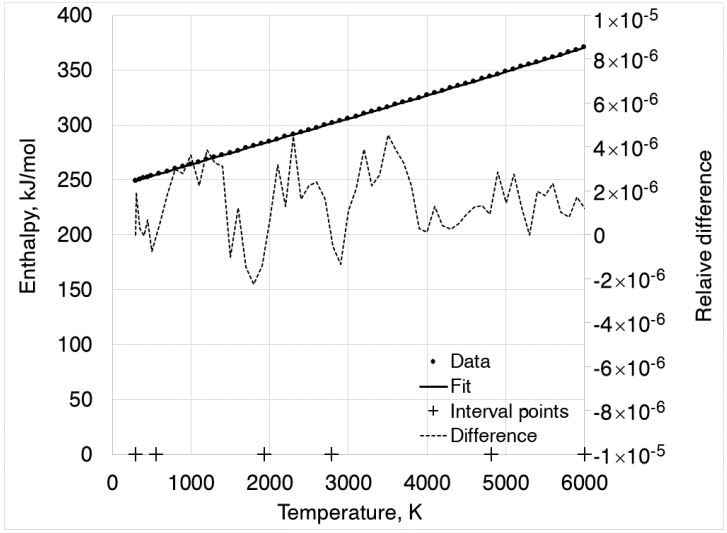


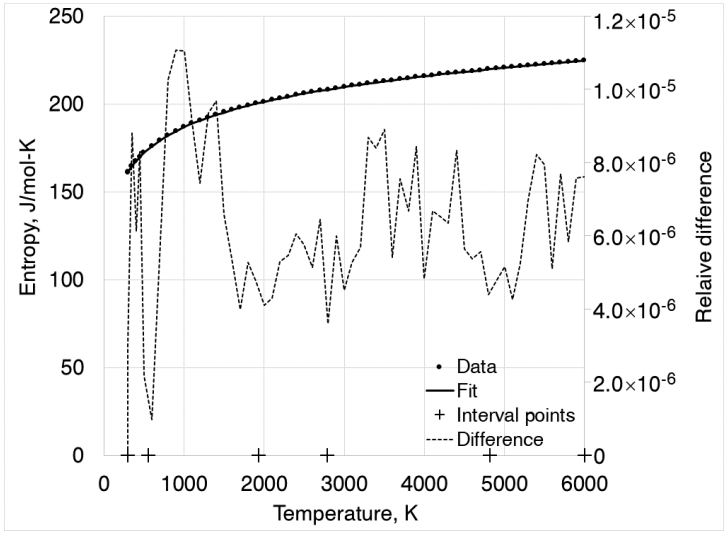


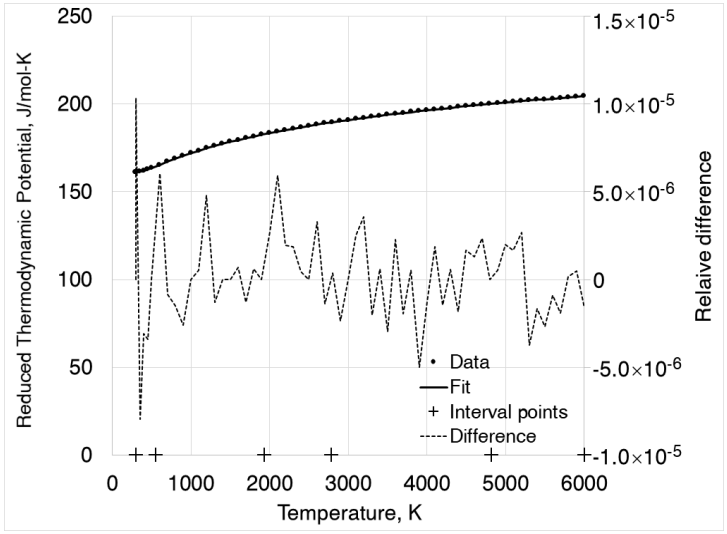
Oxygen (O) – O1(g)

MW = 15.9994 g/mol – $\Delta_f H^o$(298.15 K) = 249.173 kJ/mol – S^o(298.15 K) = 161.058 J/mol-K – p^o = 0.1 MPa



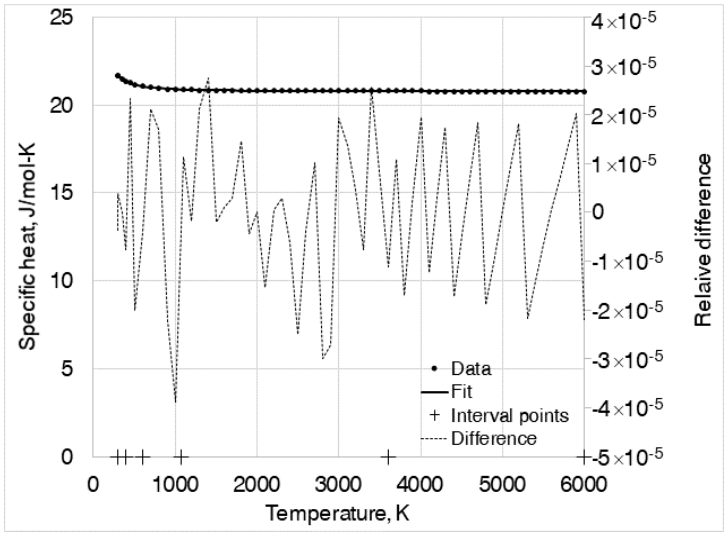


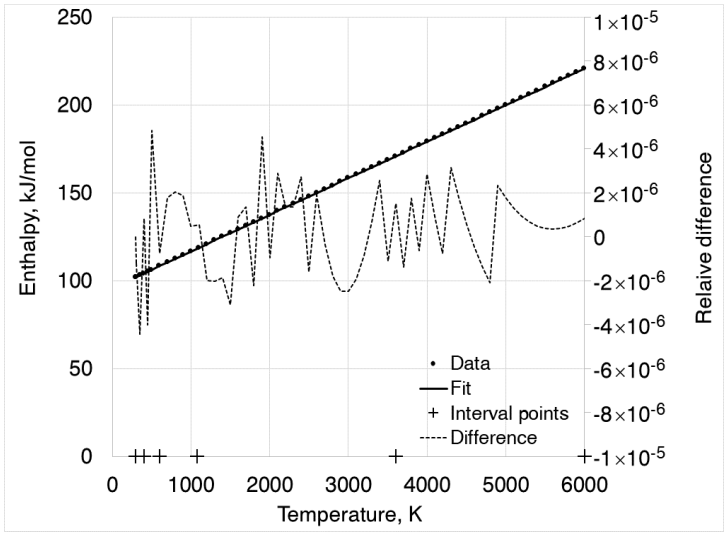


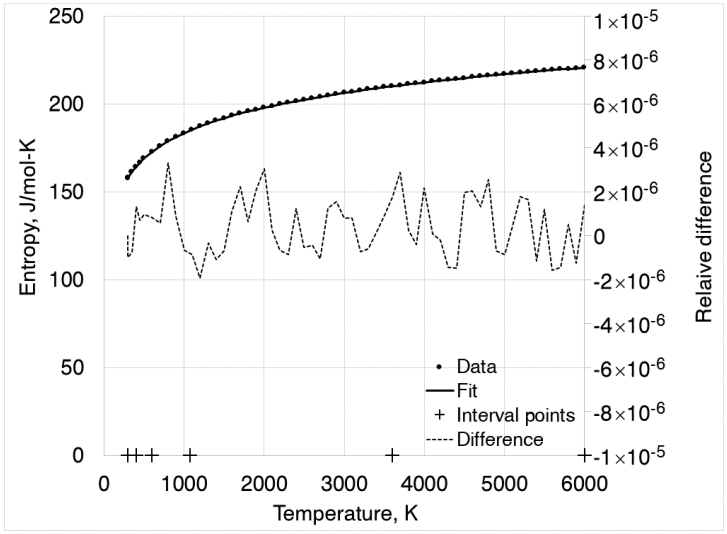


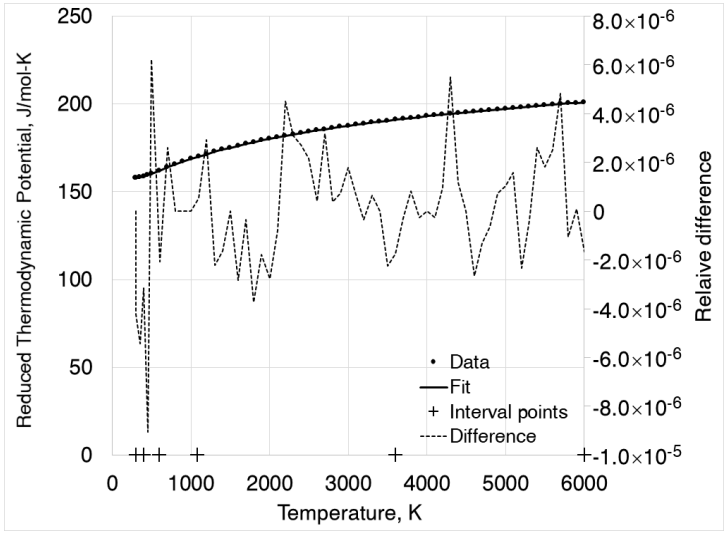
Oxygen, Ion (O^-) – O1-(g)

MW = 15.9999 g/mol – $\Delta_f H^o$(298.15 K) = 101.846 kJ/mol – S^o(298.15 K) = 157.79 J/mol-K – p^o = 0.1 MPa



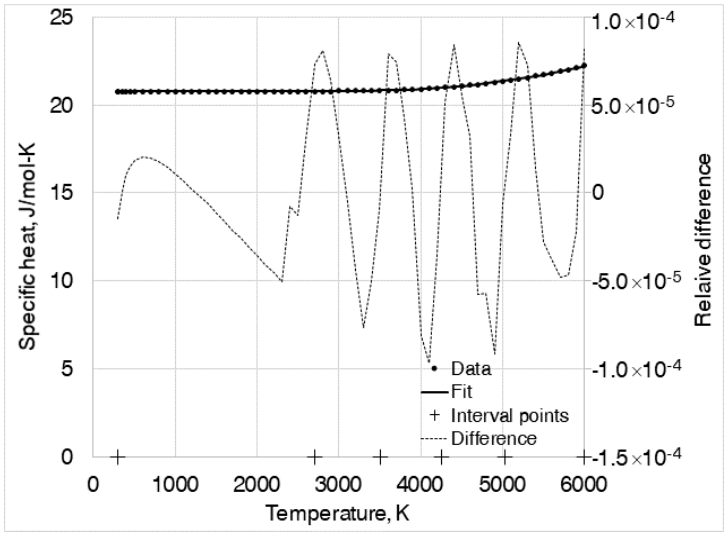


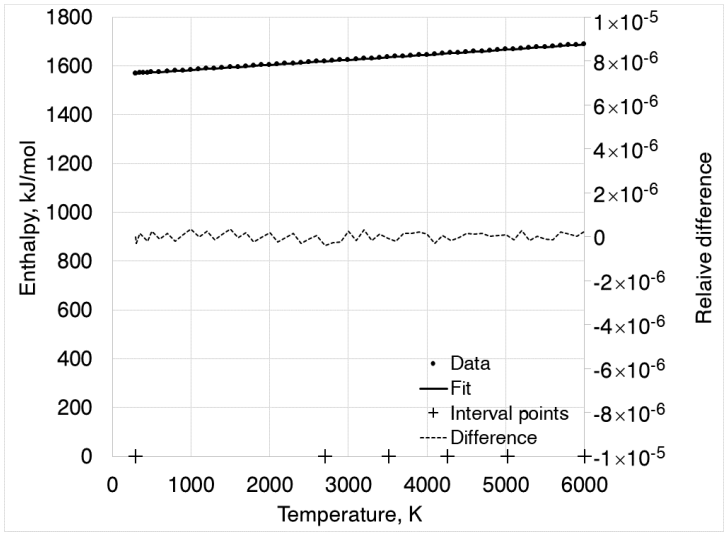


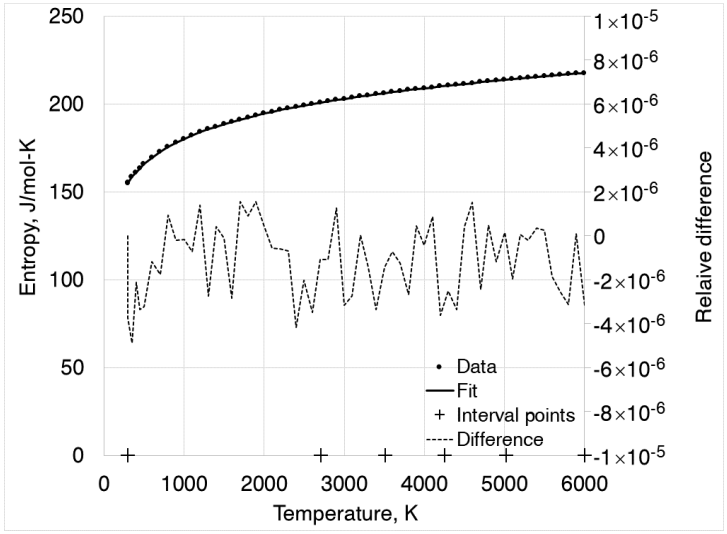


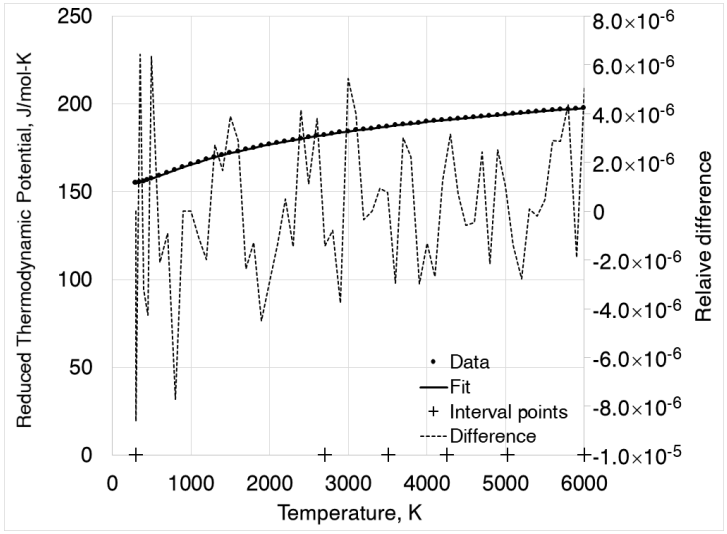
Oxygen, Ion (O^+) – O1+(g)

MW = 15.9989 g/mol – $\Delta_f H^o$(298.15 K) = 1568.786 kJ/mol – S^o(298.15 K) = 154.959 J/mol-K – p^o = 0.1 MPa



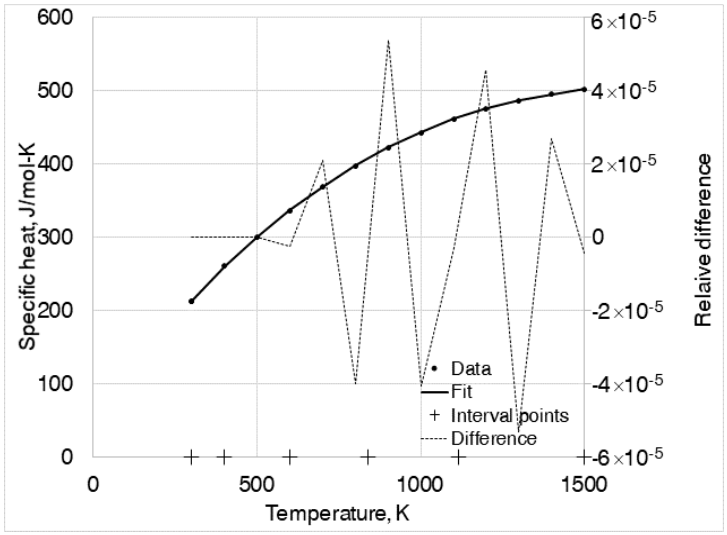


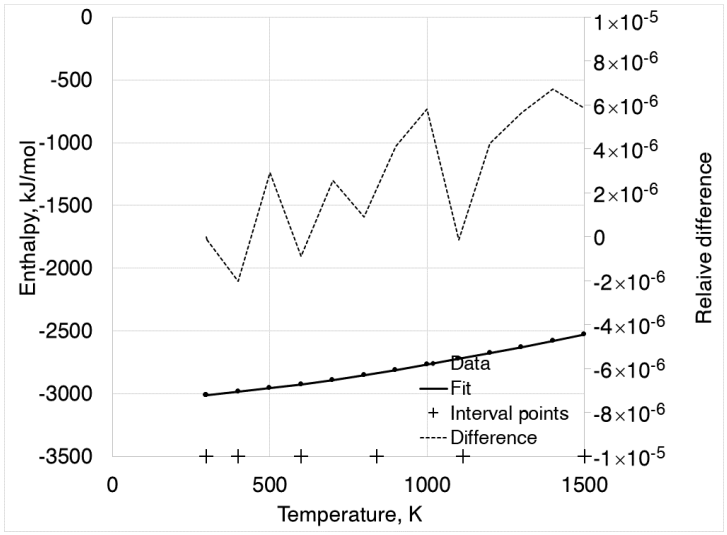


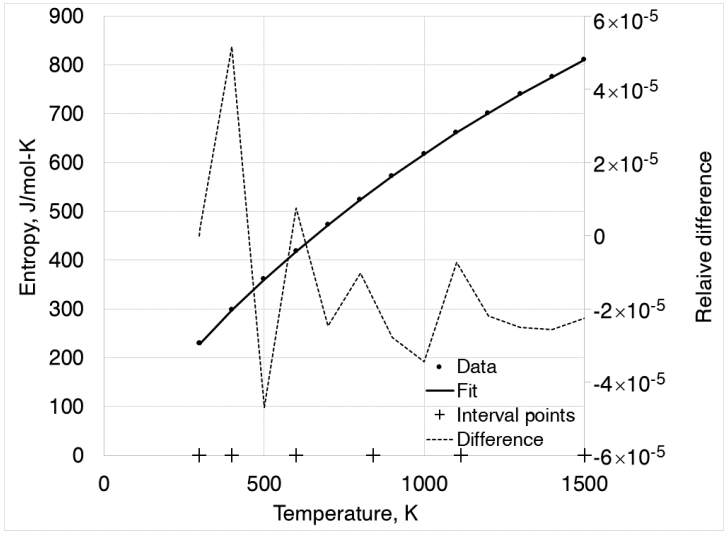


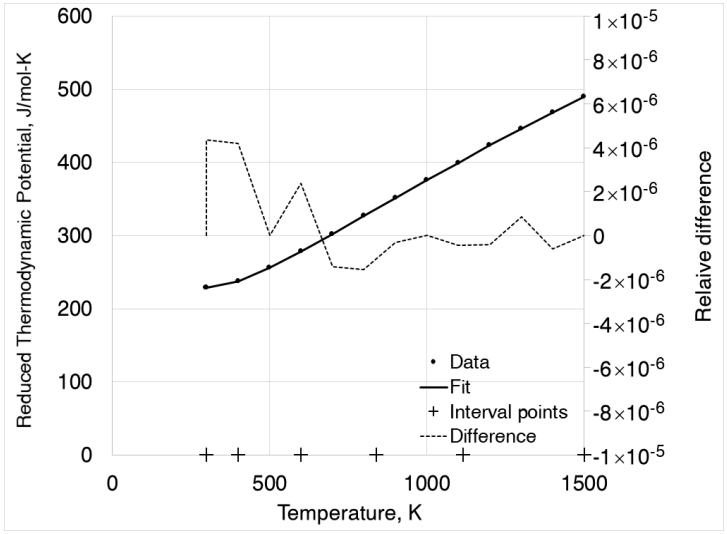
Phosphorus Oxide ($(P_2O_5)_2$) – O10P4(cr)

MW = 283.889 g/mol – $\Delta_f H^o$(298.15 K) = -3009.936 kJ/mol – S^o(298.15 K) = 228.781 J/mol-K – p^o = 0.1 MPa



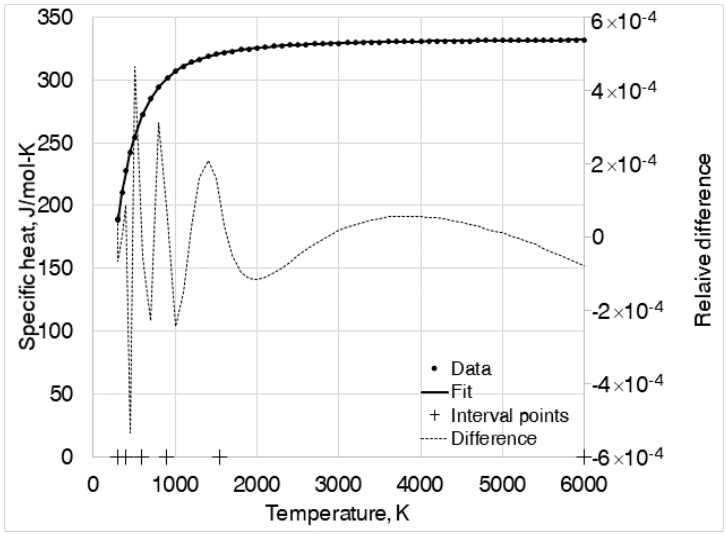


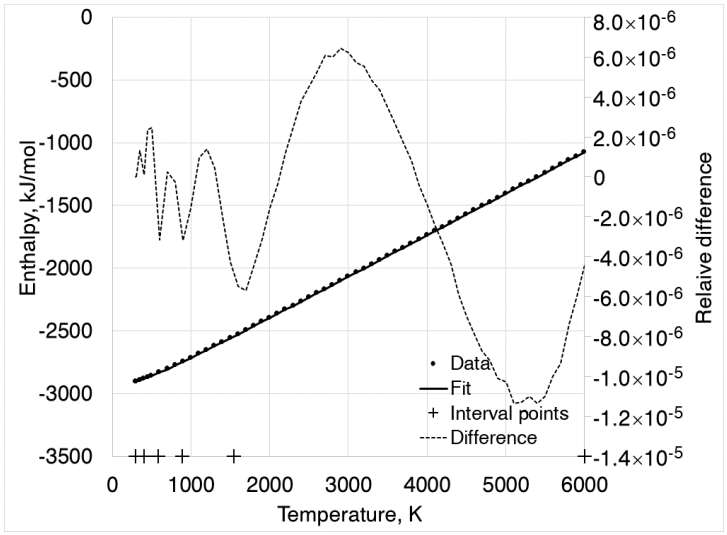


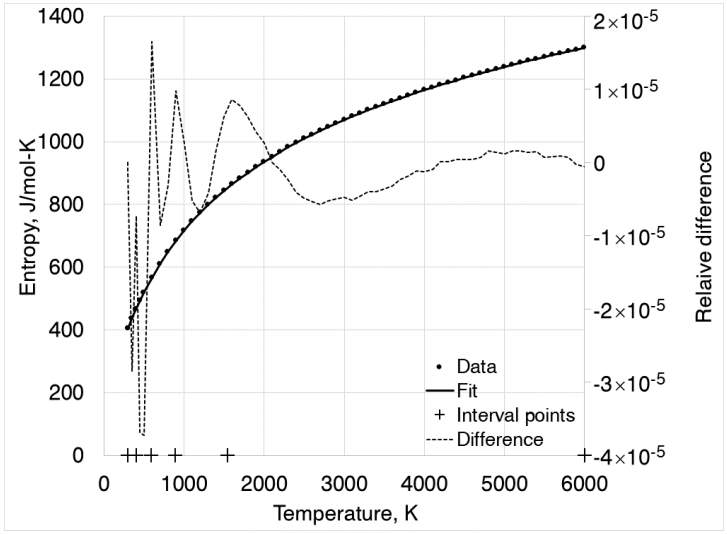


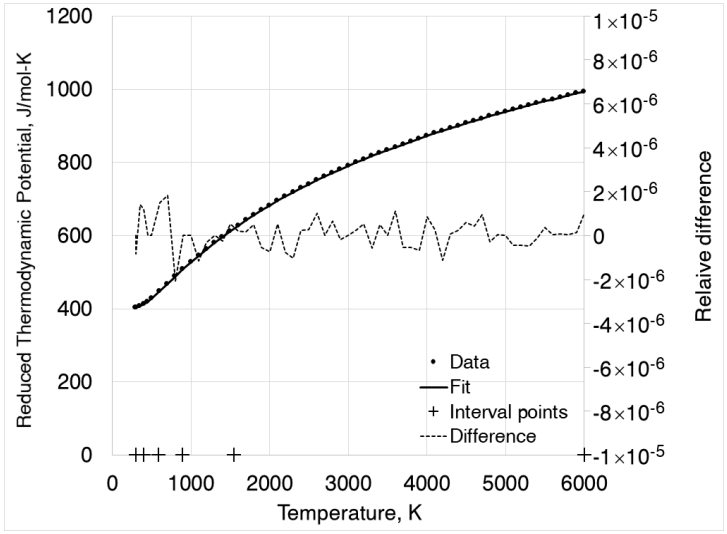
Phosphorus Oxide ($(P_2O_5)_2$) – O10P4(g)

MW = 283.889 g/mol – $\Delta_f H^o$(298.15 K) = -2904.081 kJ/mol – S^o(298.15 K) = 403.964 J/mol-K – p^o = 0.1 MPa



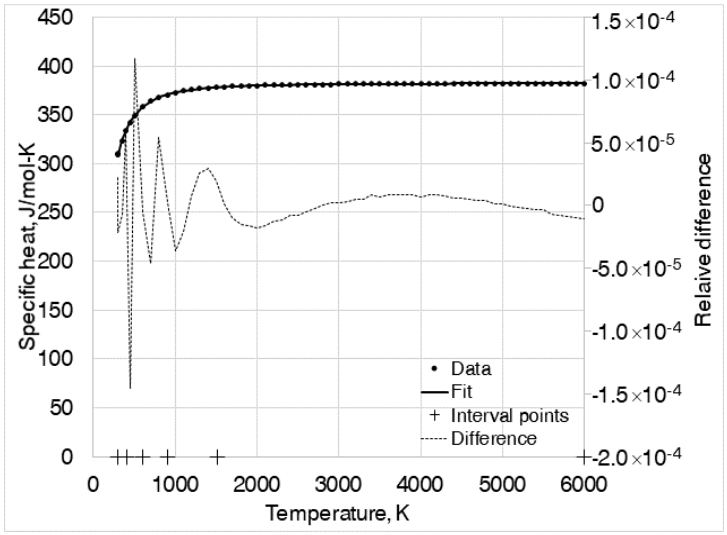


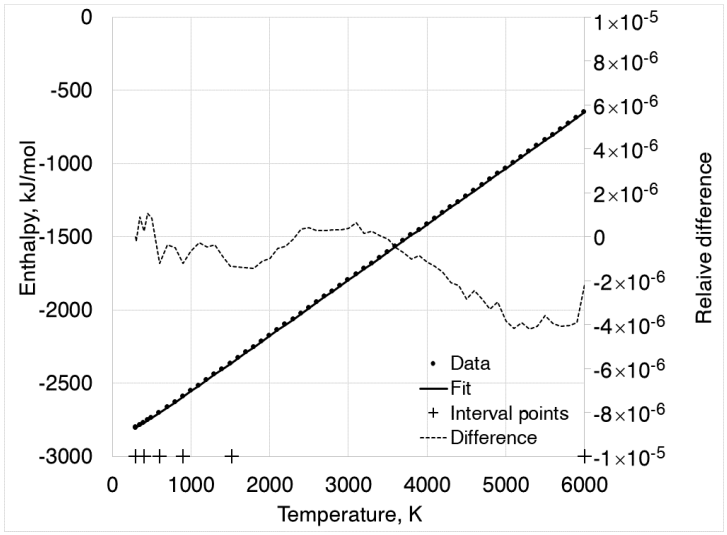


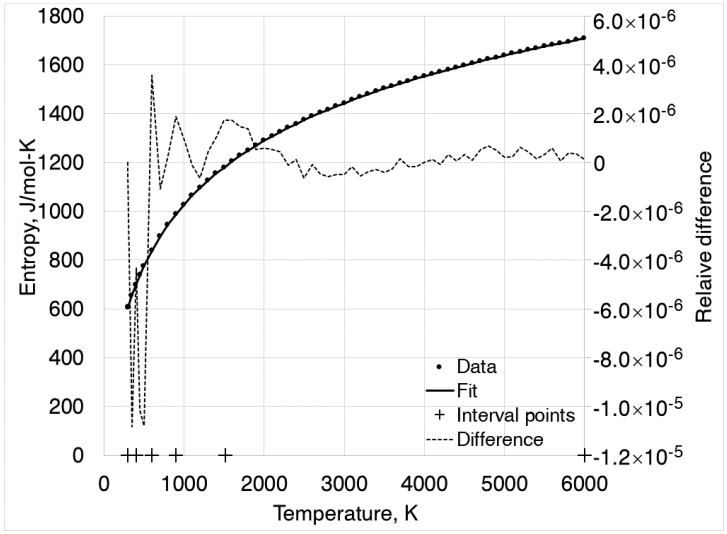


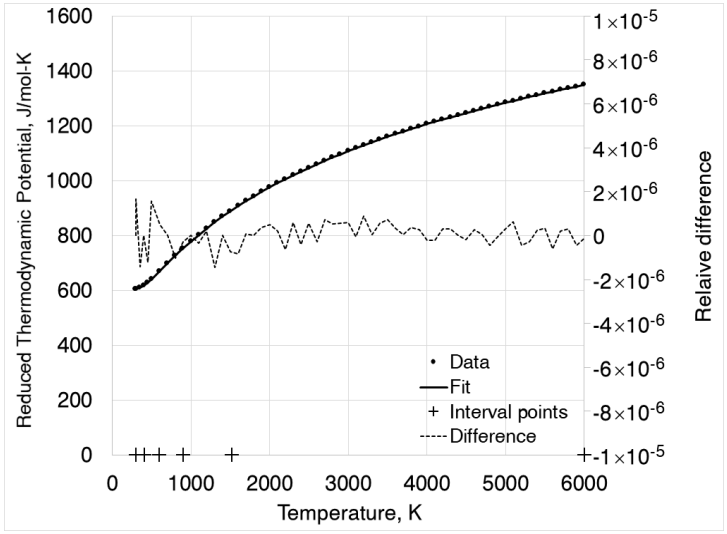
Tungsten Oxide ($(WO_3)_4$) – O12W4(g)

MW = 927.39 g/mol – $\Delta_f H^o$(298.15 K) = -2804.117 kJ/mol – S^o(298.15 K) = 605.272 J/mol-K – p^o = 0.1 MPa



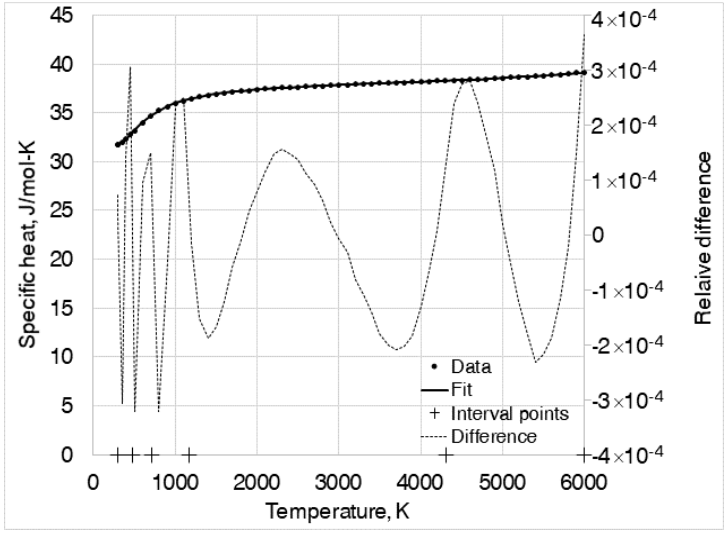


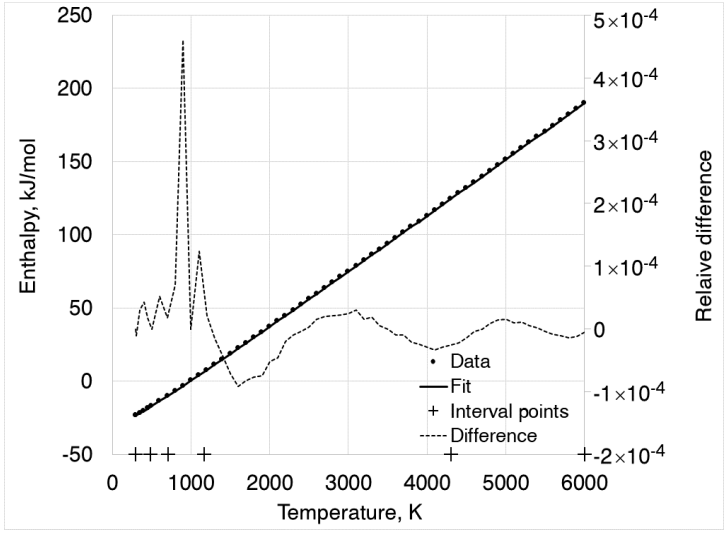


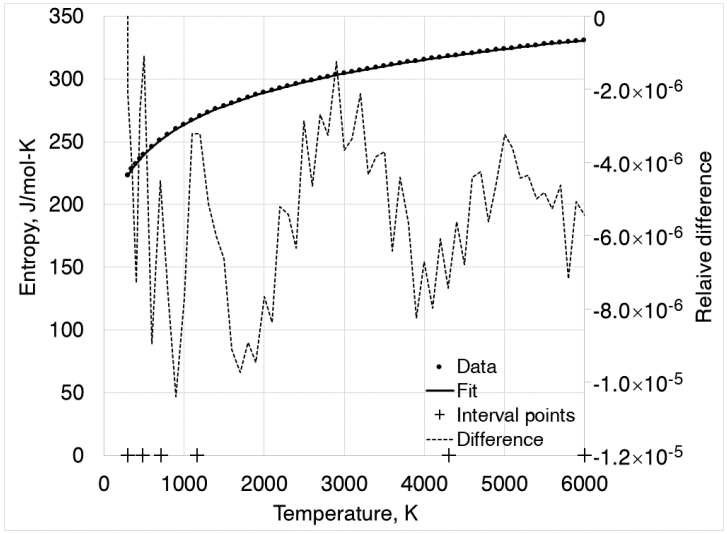


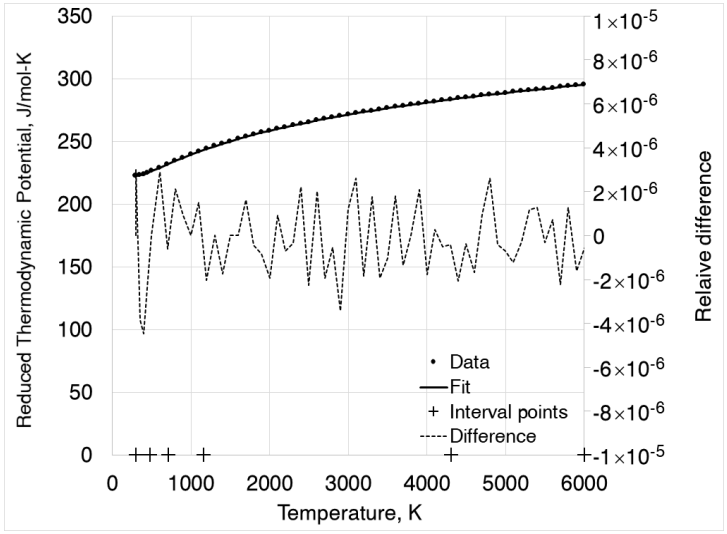
Phosphorus Oxide (PO) – O1P1(g)

MW = 46.9732 g/mol – $\Delta_f H^o$(298.15 K) = -23.548 kJ/mol – S^o(298.15 K) = 222.776 J/mol-K – p^o = 0.1 MPa



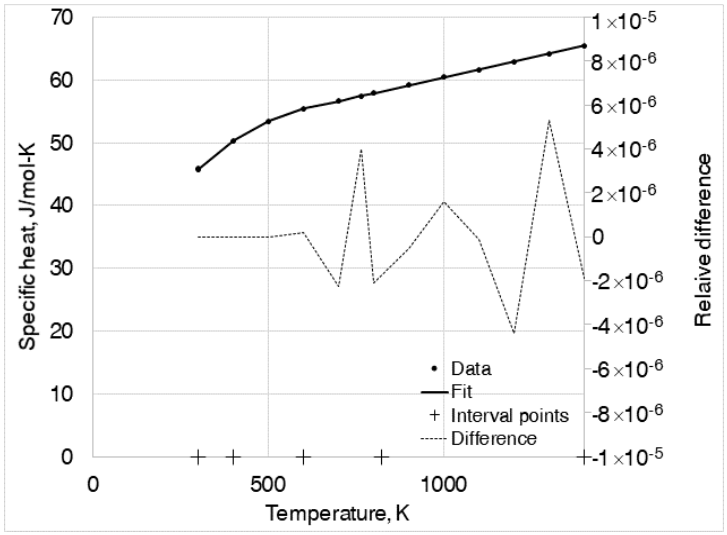


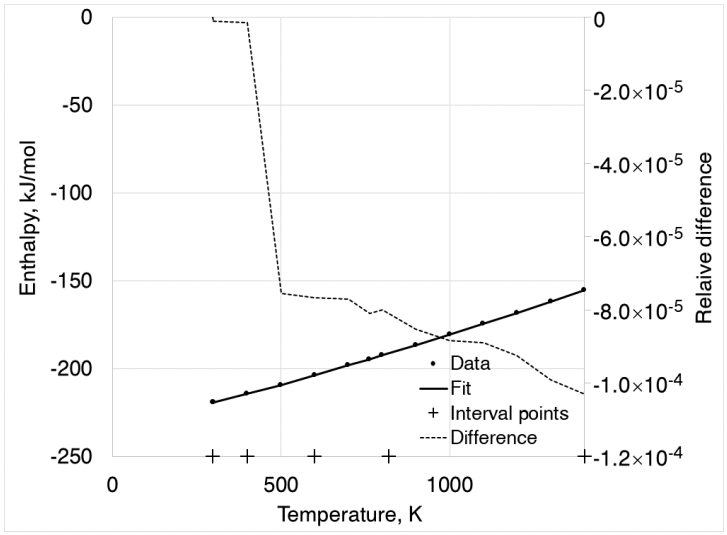


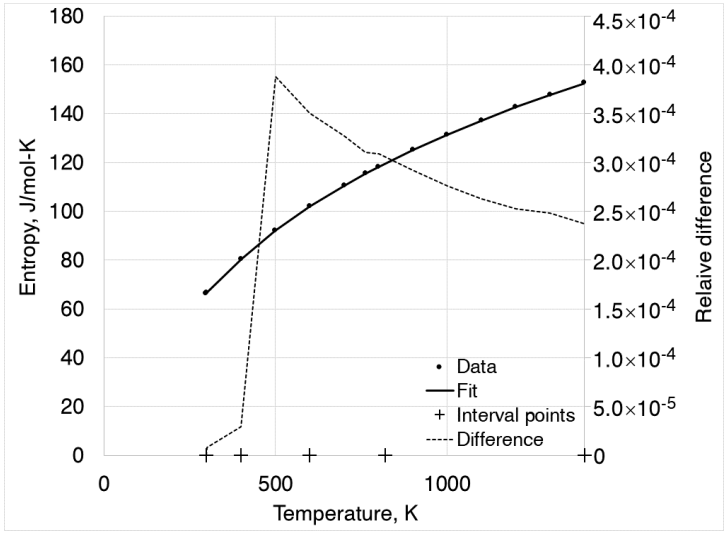


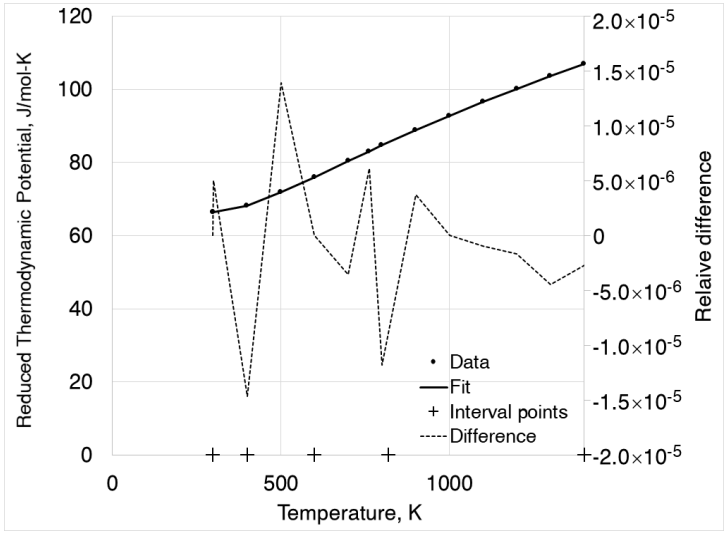
Lead Oxide, Red (PbO) – O1Pb1(cr)_01

MW = 223.2 g/mol – $\Delta_f H^o$(298.15 K) = -219.409 kJ/mol – S^o(298.15 K) = 66.316 J/mol-K – p^o = 0.1 MPa



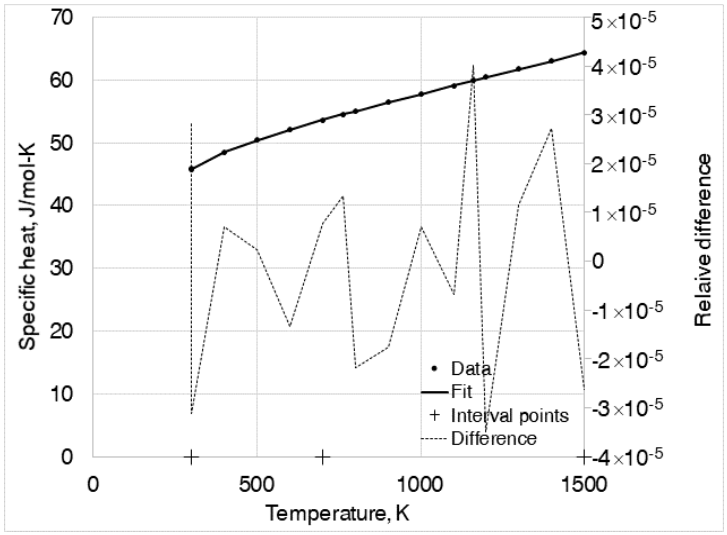


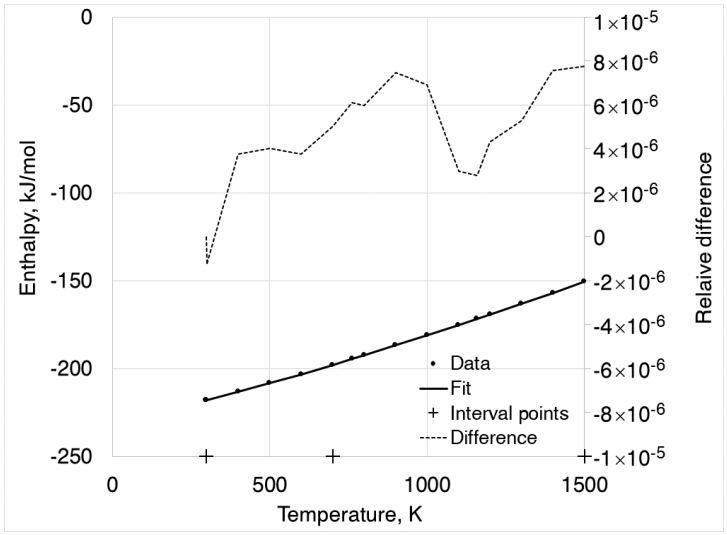


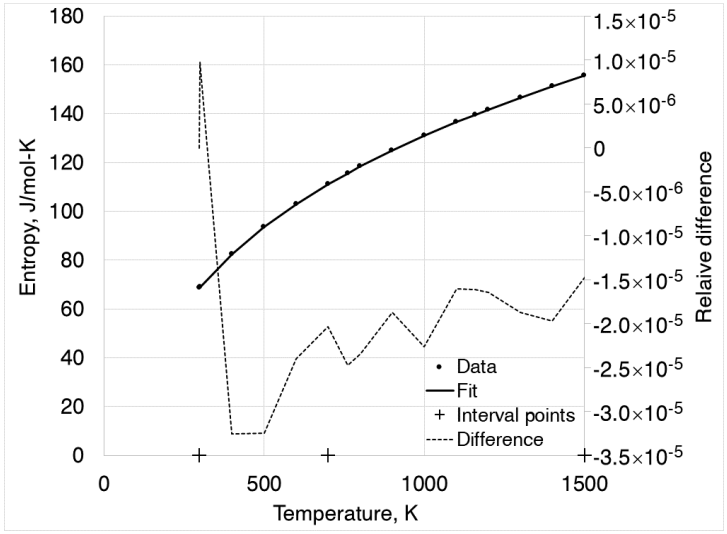


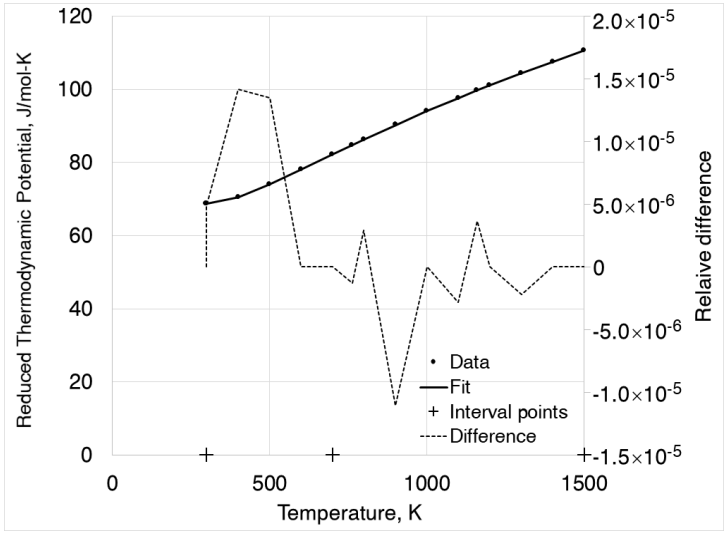
Lead Oxide, Yellow (PbO) – O1Pb1(cr)_02

MW = 223.2 g/mol – $\Delta_f H^o$(298.15 K) = -218.062 kJ/mol – S^o(298.15 K) = 68.701 J/mol-K – p^o = 0.1 MPa



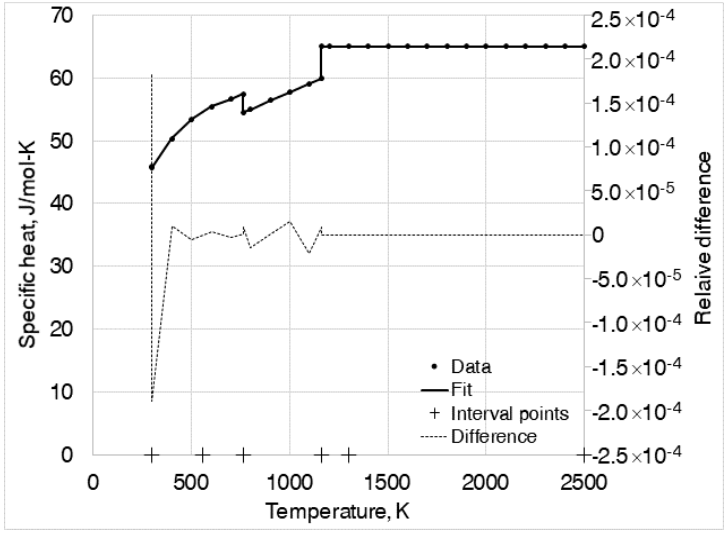


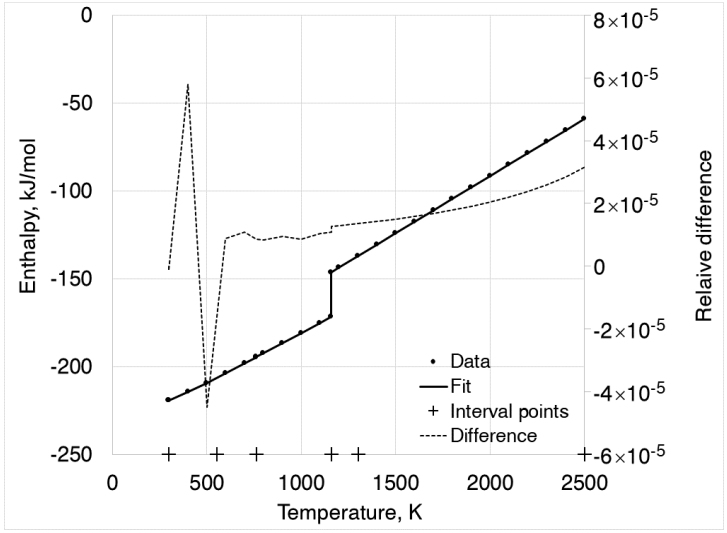


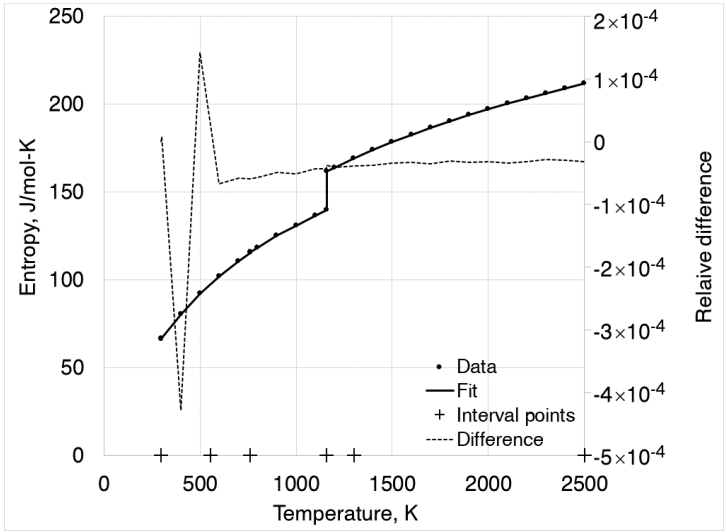


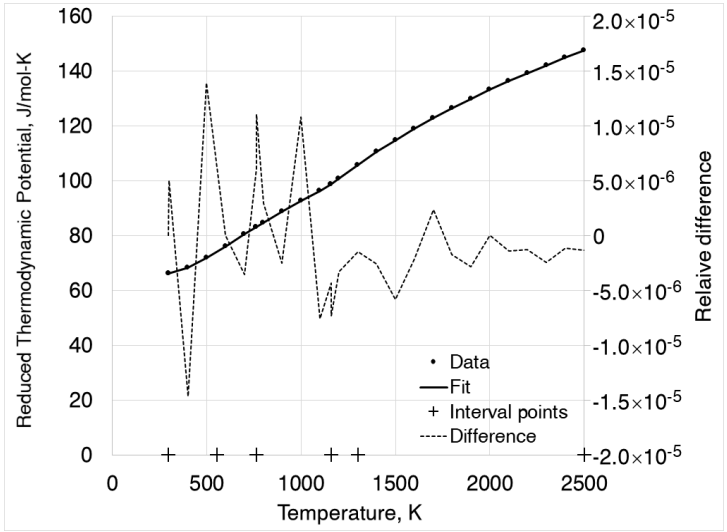
Lead Oxide (PbO) – O1Pb1(cr,l)

MW = 223.2 g/mol – $\Delta_f H^o$(298.15 K) = -219.409 kJ/mol – S^o(298.15 K) = 66.316 J/mol-K – p^o = 0.1 MPa



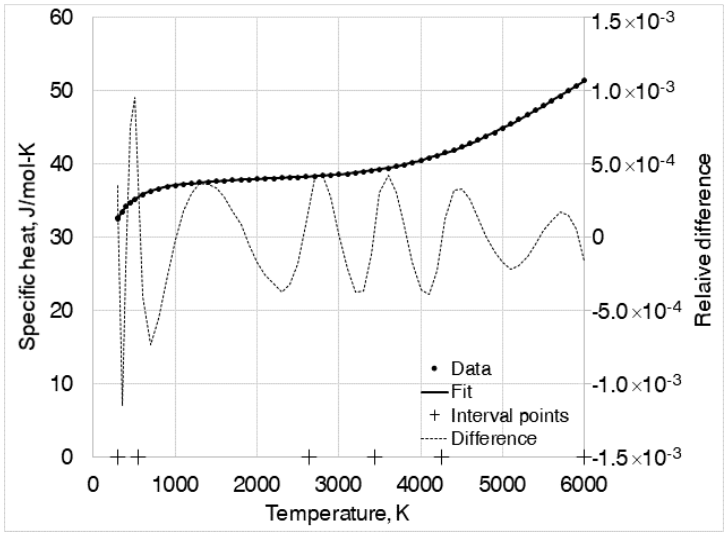


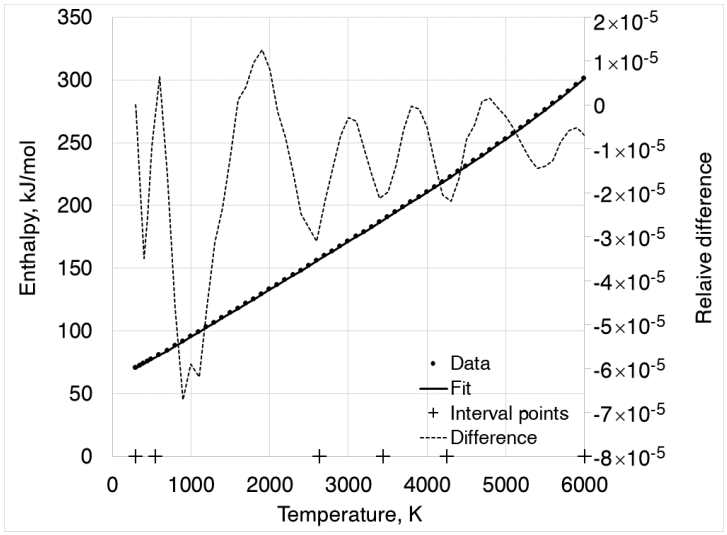


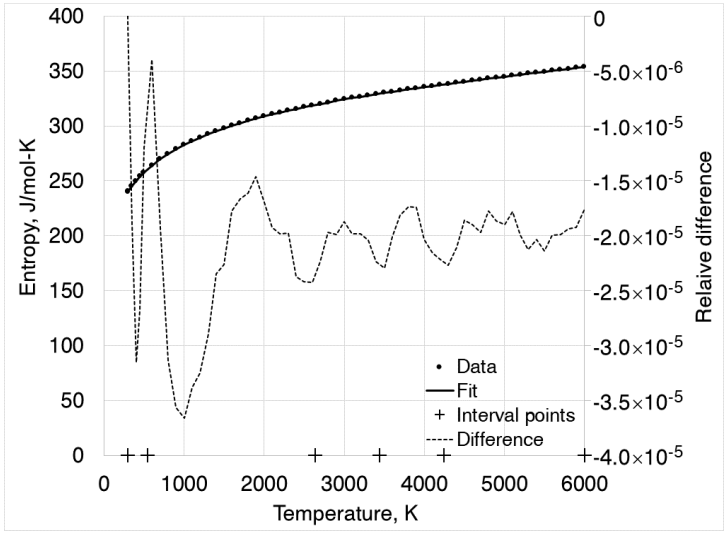


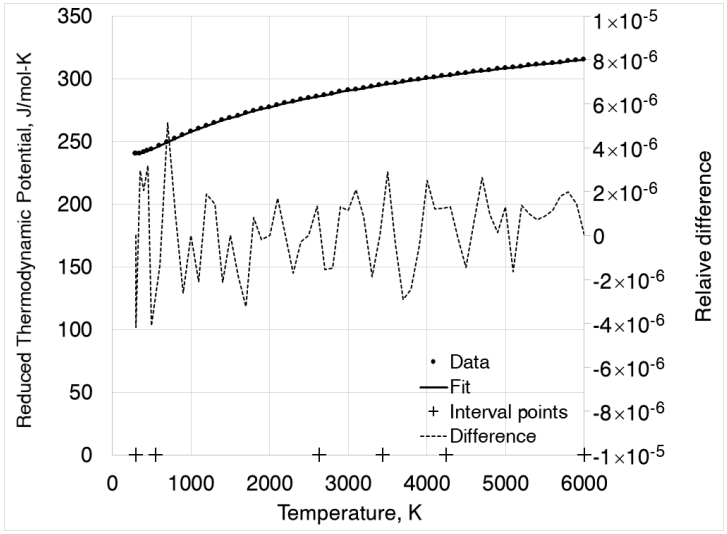
Lead Oxide (PbO) – O1Pb1(g)

MW = 223.2 g/mol – $\Delta_f H^o$(298.15 K) = 70.291 kJ/mol – S^o(298.15 K) = 240.039 J/mol-K – p^o = 0.1 MPa



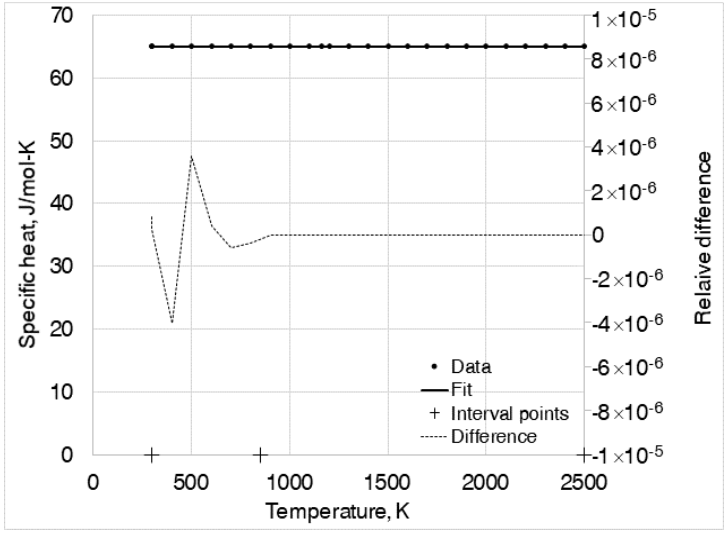


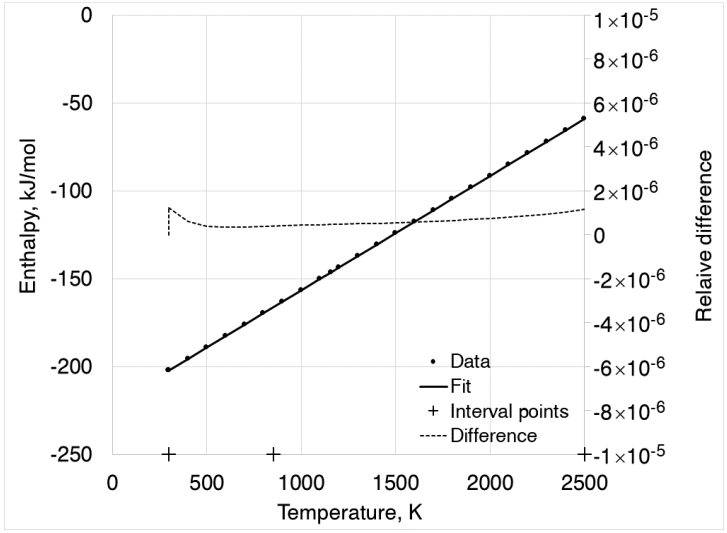


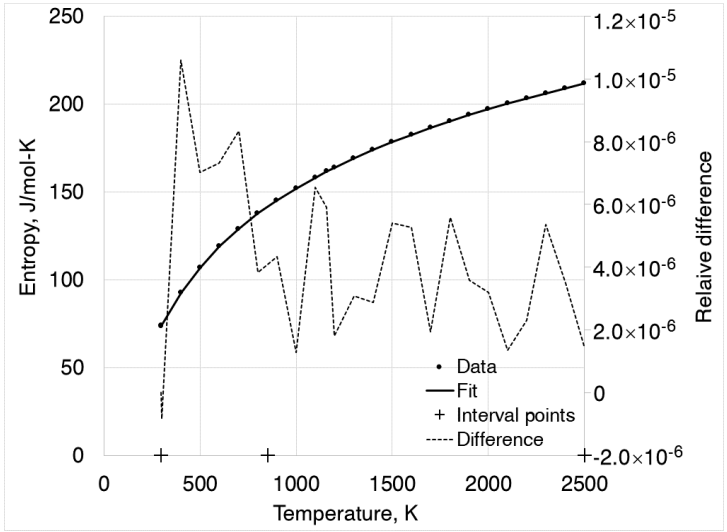


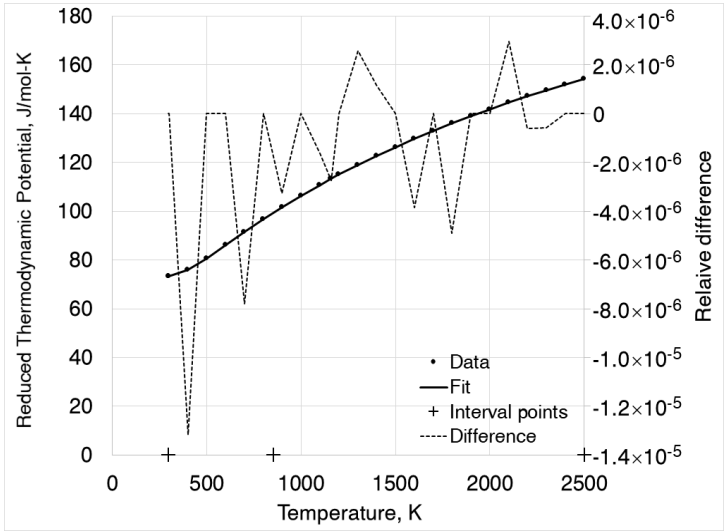
Lead Oxide (PbO) – O1Pb1(l)

MW = 223.2 g/mol – $\Delta_f H^o$(298.15 K) = -202.249 kJ/mol – S^o(298.15 K) = 73.376 J/mol-K – p^o = 0.1 MPa



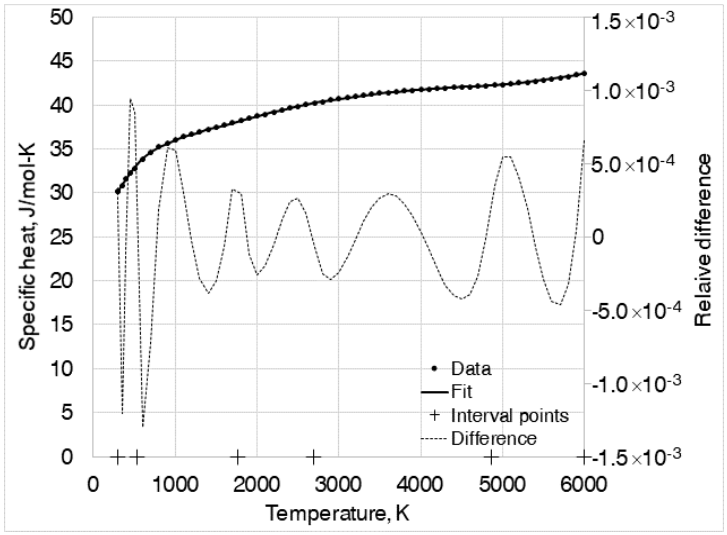


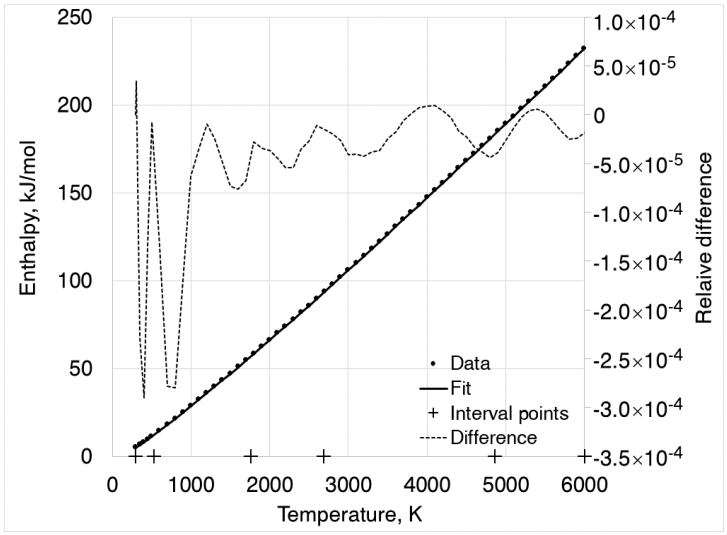


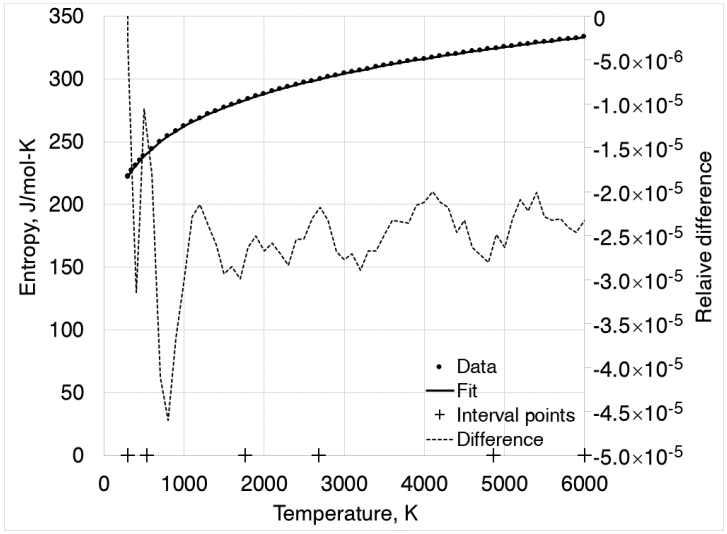


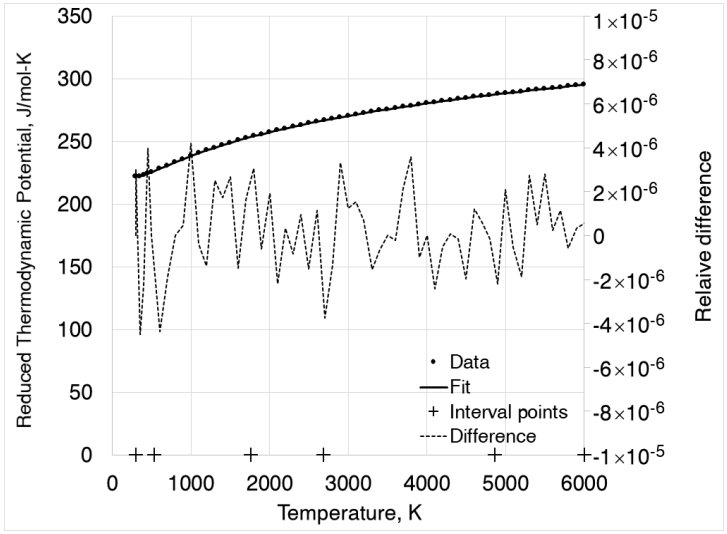
Sulfur Oxide (SO) – O1S1(g)

MW = 48.06 g/mol – $\Delta_f H^o$(298.15 K) = 5.007 kJ/mol – S^o(298.15 K) = 221.944 J/mol-K – p^o = 0.1 MPa



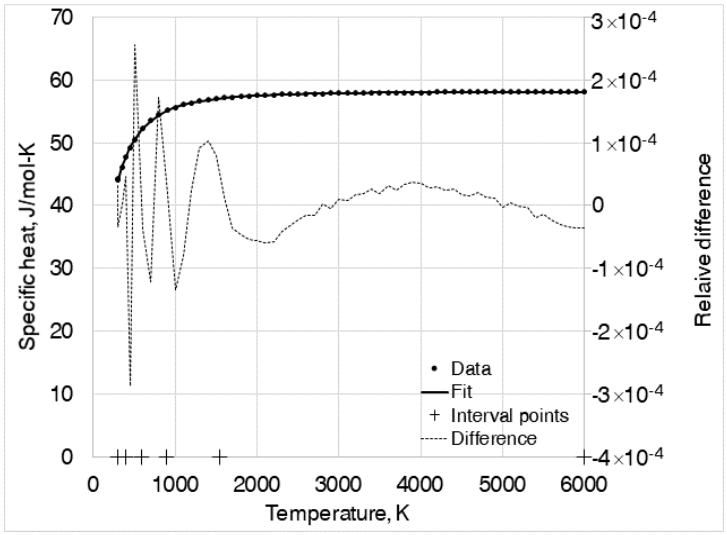


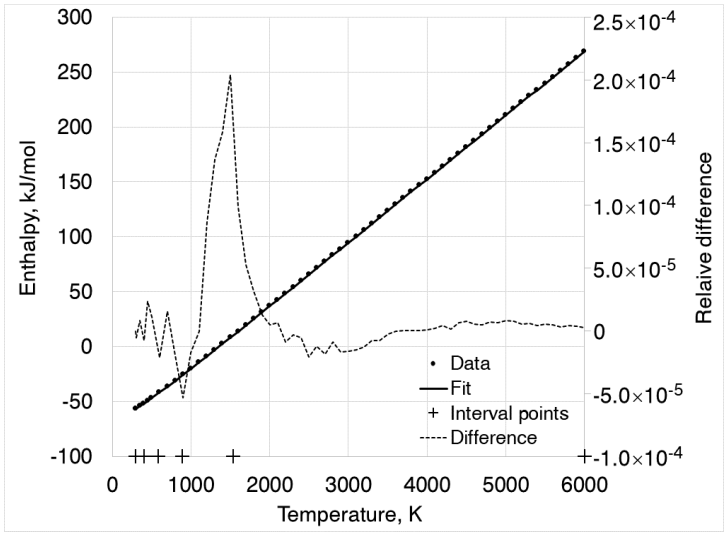


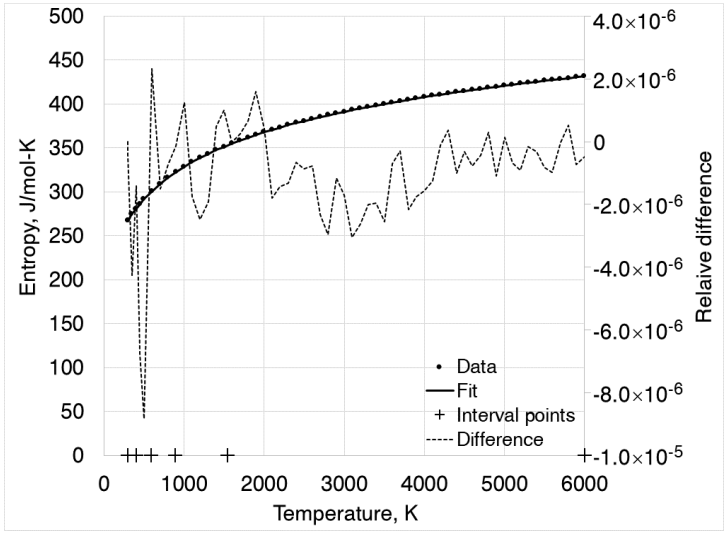


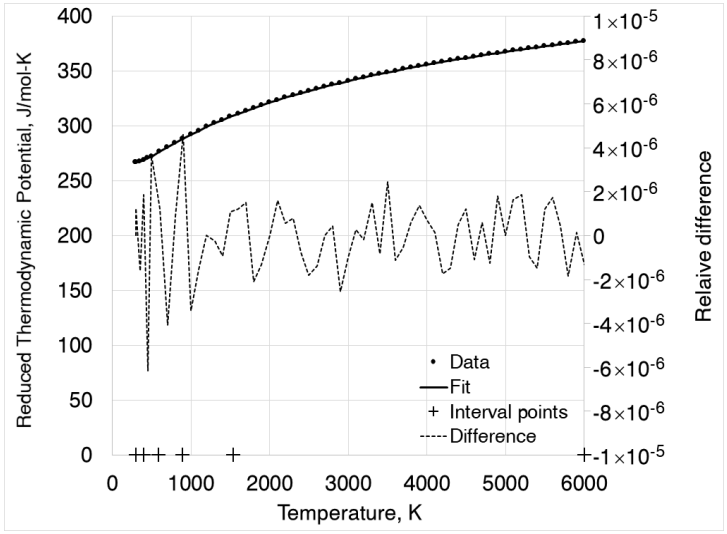
Sulfur Oxide (SSO) – O1S2(g)

MW = 80.12 g/mol – $\Delta_f H^o$(298.15 K) = -56.484 kJ/mol – S^o(298.15 K) = 267.02 J/mol-K – p^o = 0.1 MPa



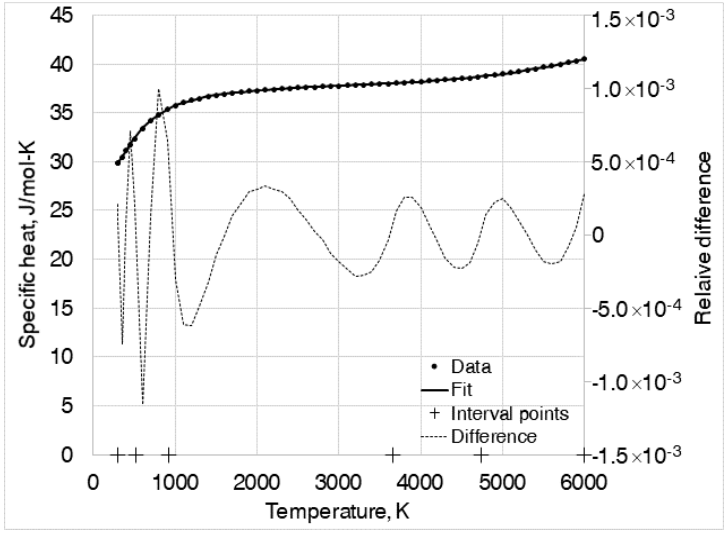


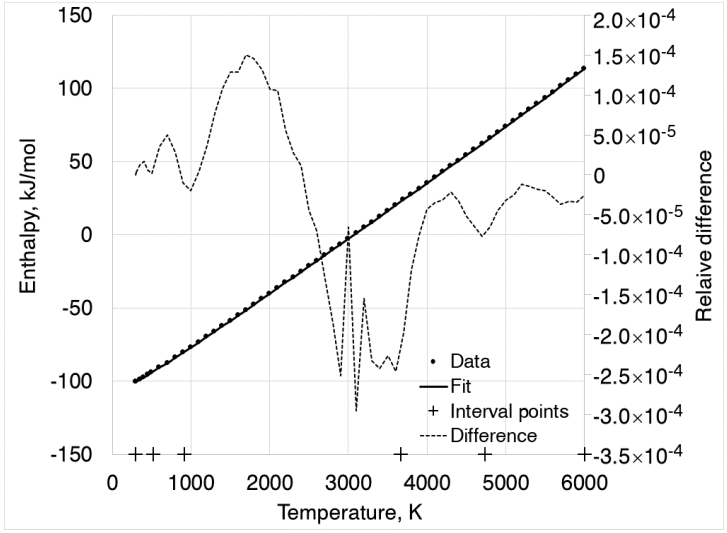


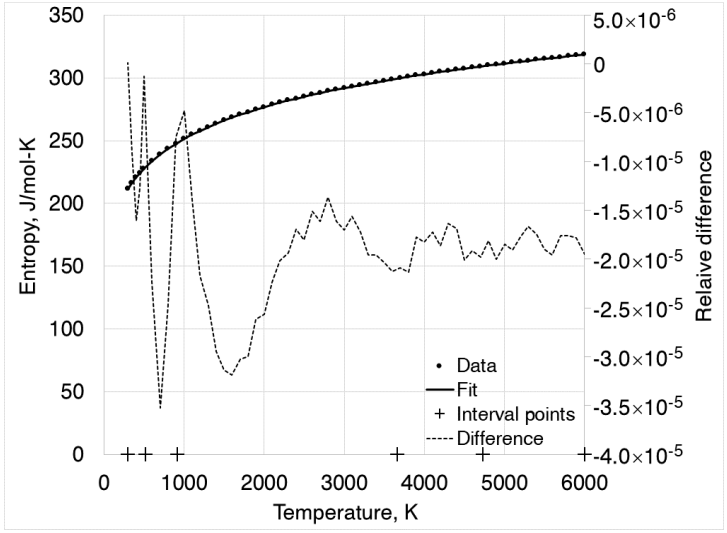


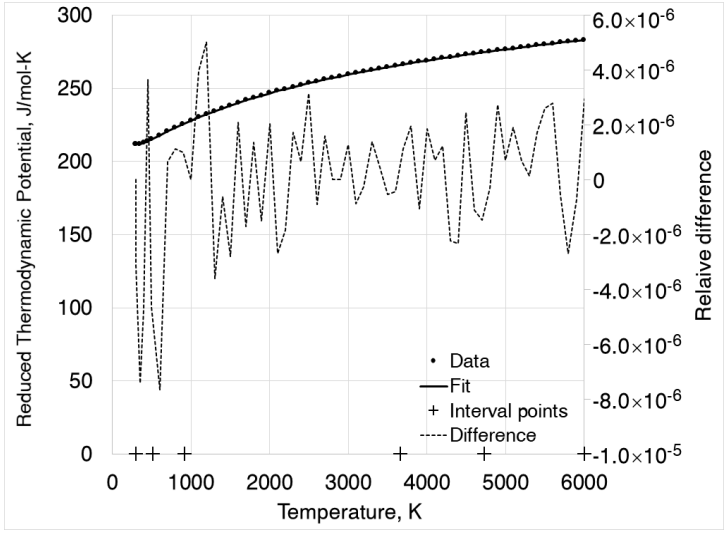
Silicon Oxide (SiO) – O1Si1(g)

MW = 44.0849 g/mol – $\Delta_f H^o$(298.15 K) = -100.416 kJ/mol – S^o(298.15 K) = 211.579 J/mol-K – p^o = 0.1 MPa



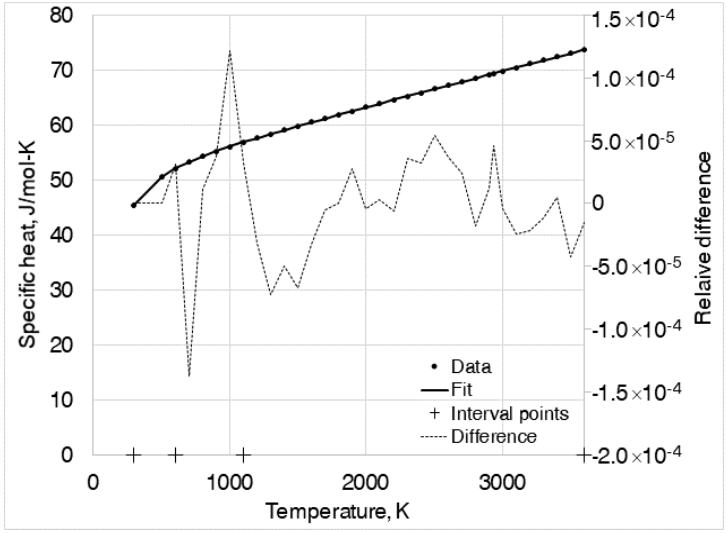


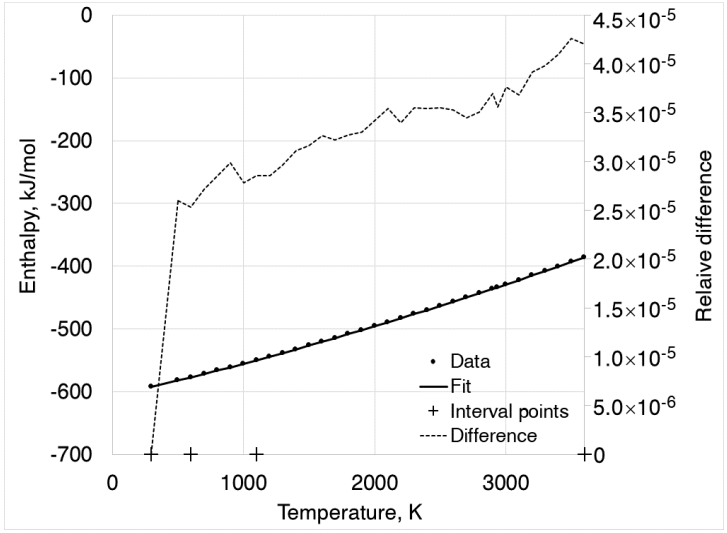


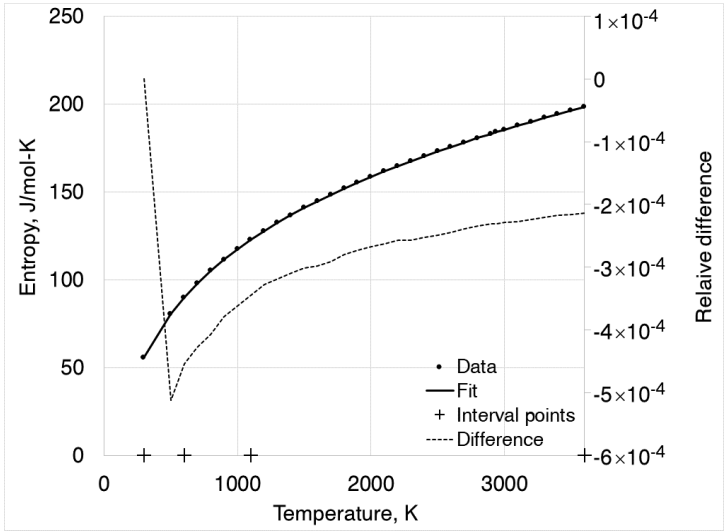


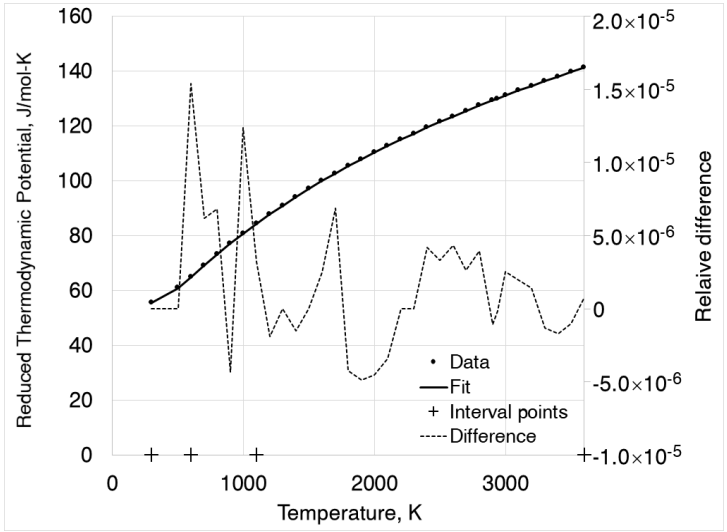
Strontium Oxide (SrO) – O1Sr1(cr)

MW = 103.62 g/mol – $\Delta_f H^o$(298.15 K) = -592.036 kJ/mol – S^o(298.15 K) = 55.522 J/mol-K – p^o = 0.1 MPa



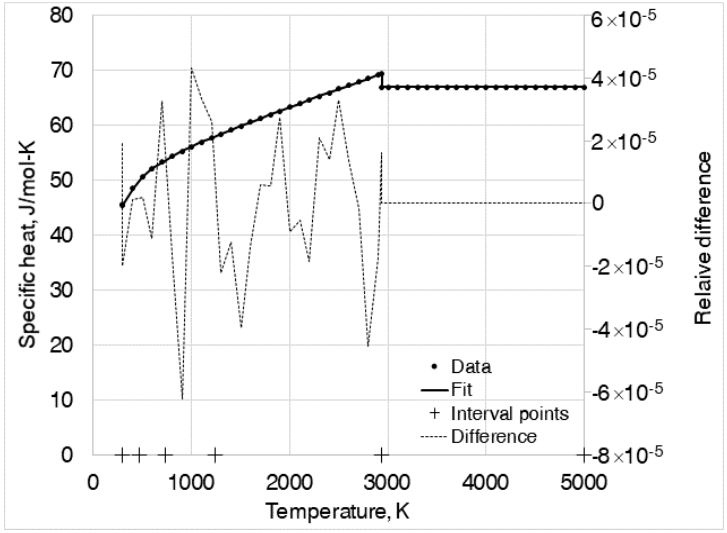


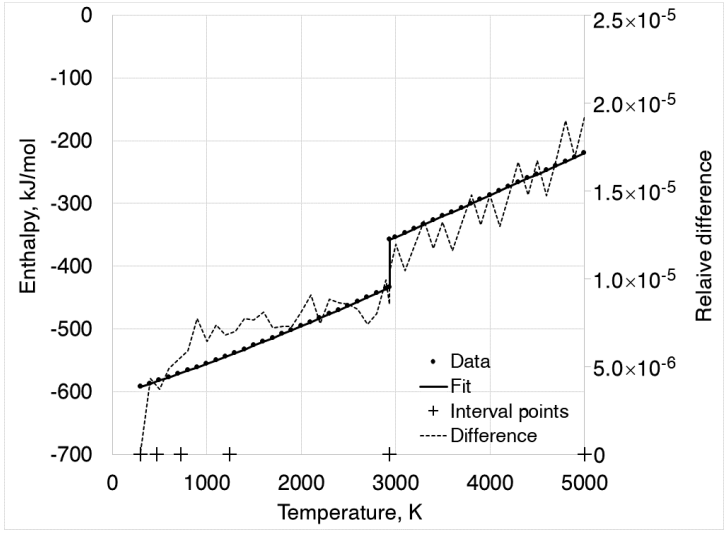


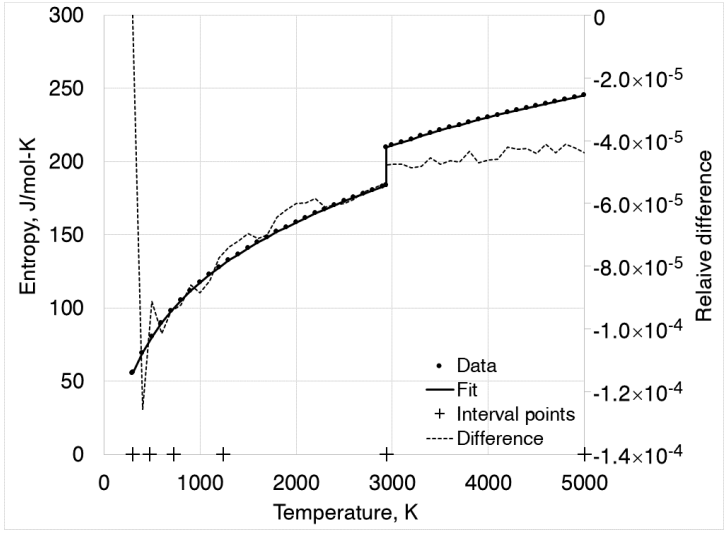


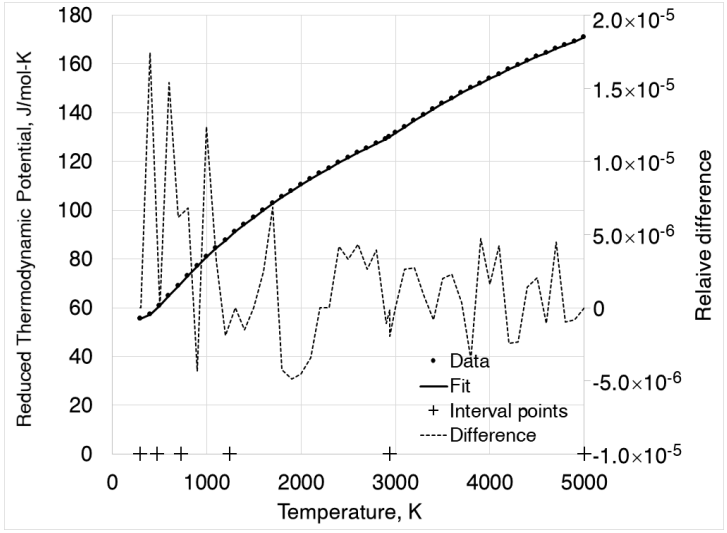
Strontium Oxide (SrO) – O1Sr1(cr,l)

MW = 103.62 g/mol – $\Delta_f H^o$(298.15 K) = -592.036 kJ/mol – S^o(298.15 K) = 55.522 J/mol-K – p^o = 0.1 MPa



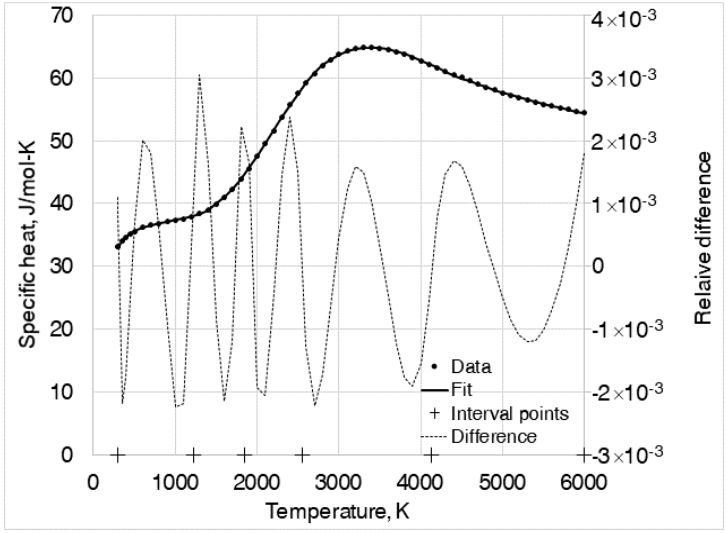


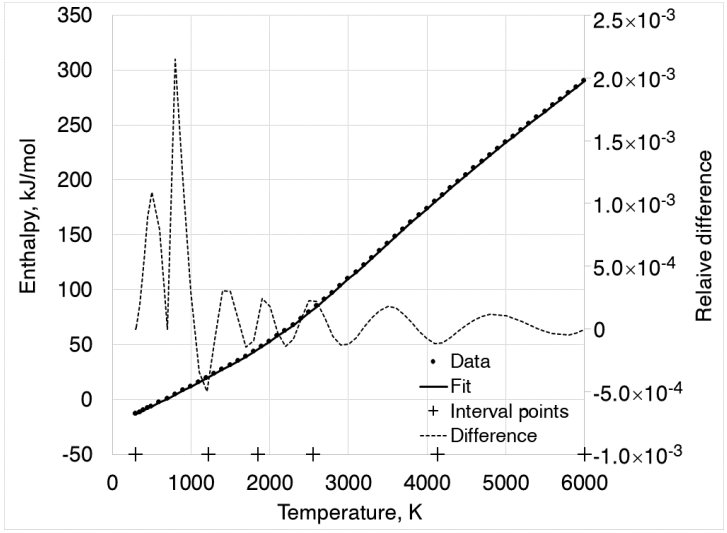


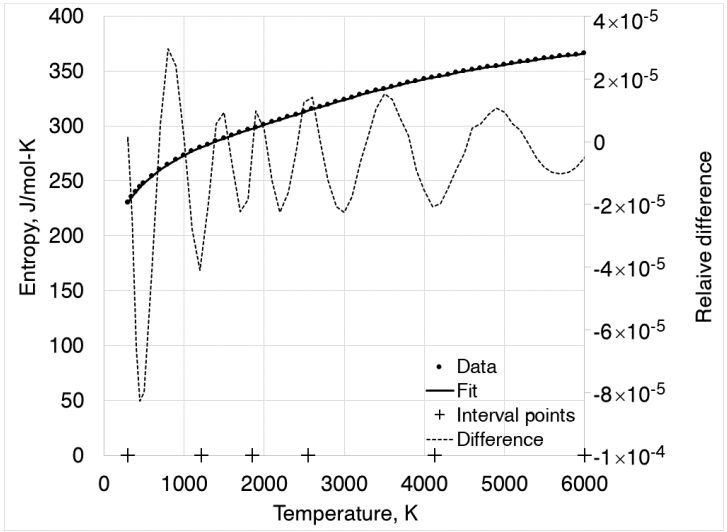


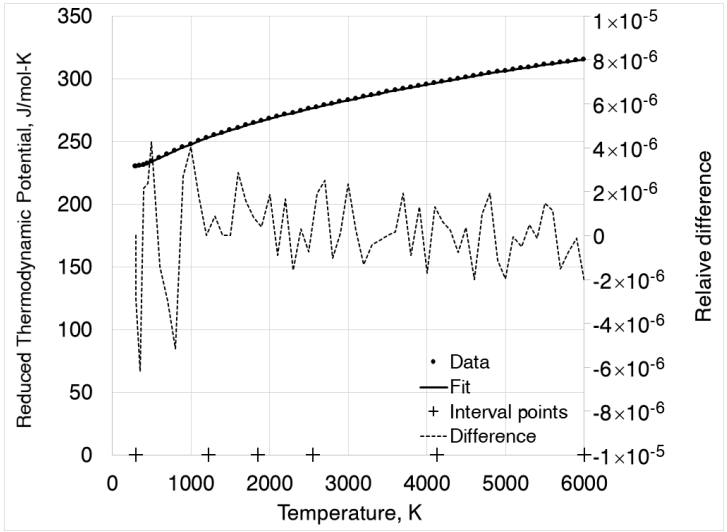
Strontium Oxide (SrO) – O1Sr1(g)

MW = 103.62 g/mol – $\Delta_f H^o$(298.15 K) = -13.389 kJ/mol – S^o(298.15 K) = 230.058 J/mol-K – p^o = 0.1 MPa



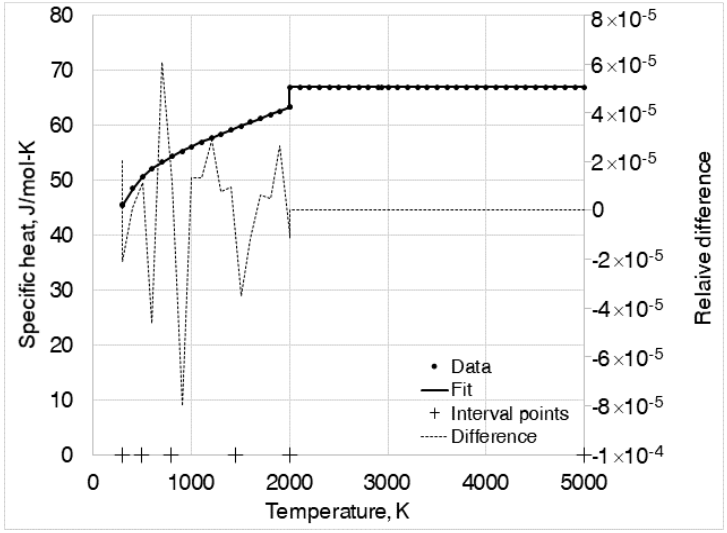


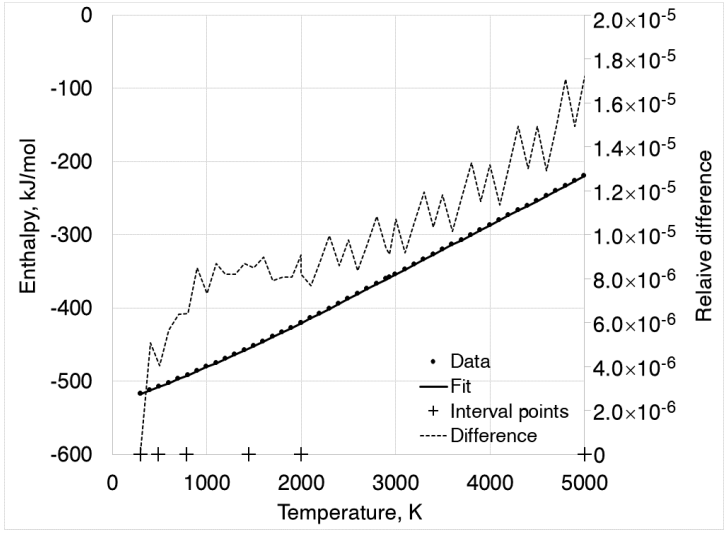


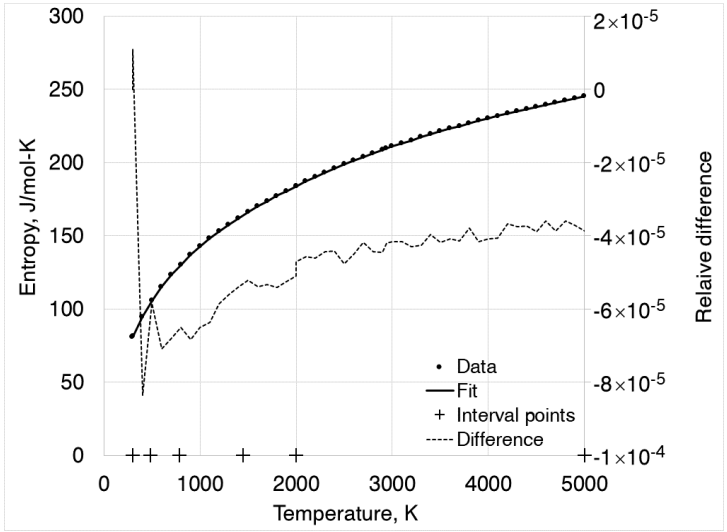


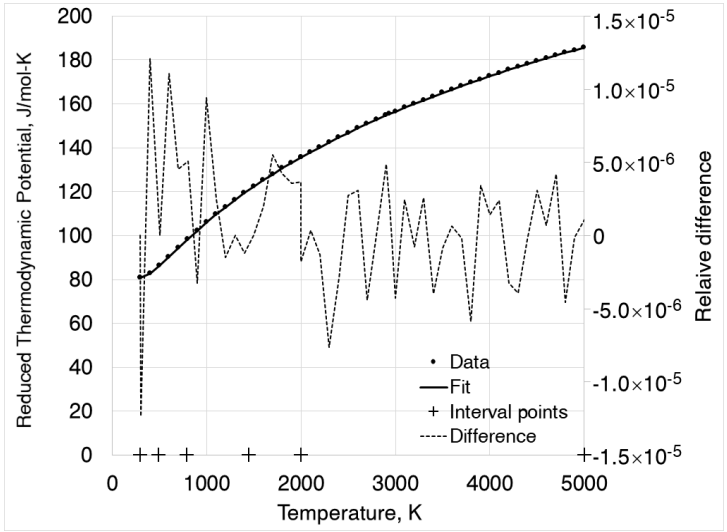
Strontium Oxide (SrO) – O1Sr1(l)

MW = 103.62 g/mol – $\Delta_f H^o$(298.15 K) = -517.292 kJ/mol – S^o(298.15 K) = 80.846 J/mol-K – p^o = 0.1 MPa



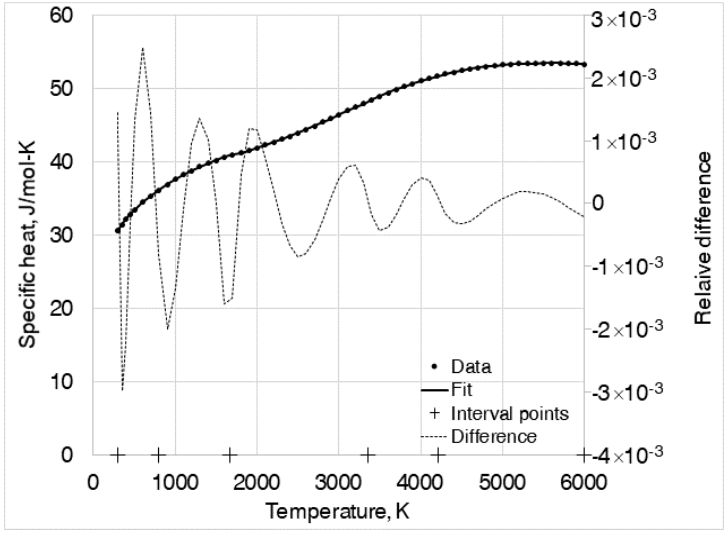


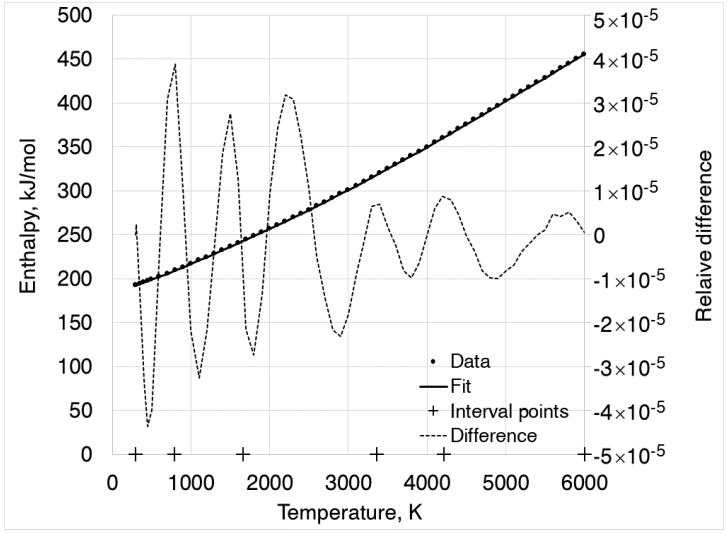


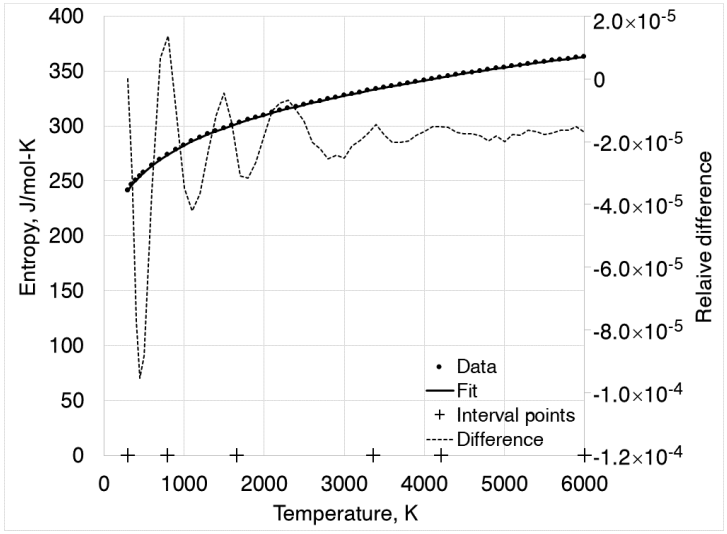


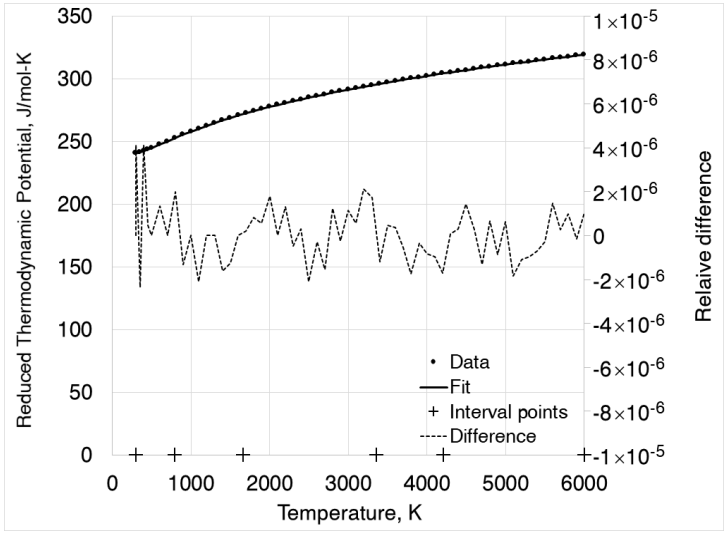
Tantalum Oxide (TaO) – O1Ta1(g)

MW = 196.9473 g/mol – $\Delta_f H^o$(298.15 K) = 192.464 kJ/mol – S^o(298.15 K) = 241.116 J/mol-K – p^o = 0.1 MPa



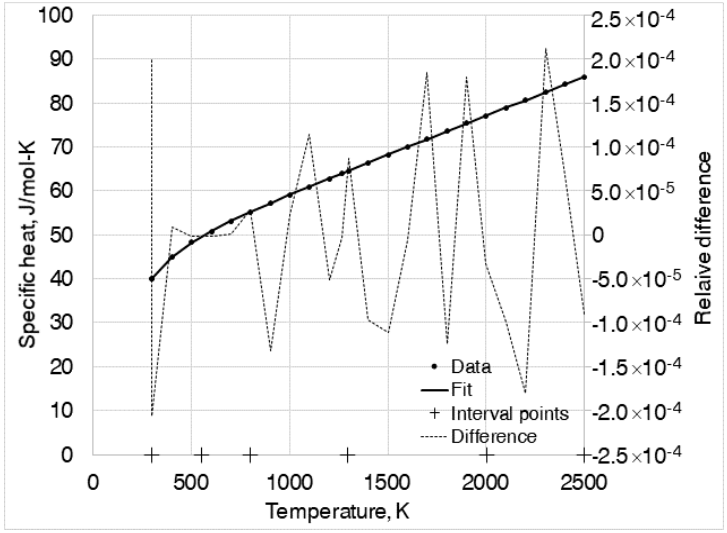


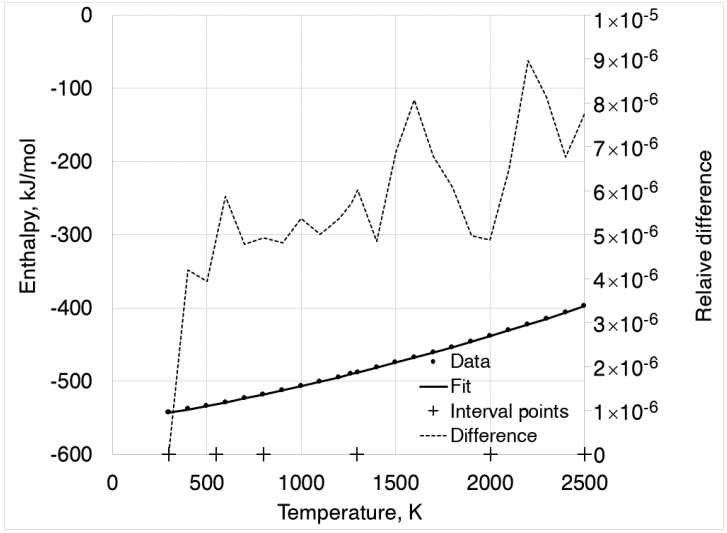


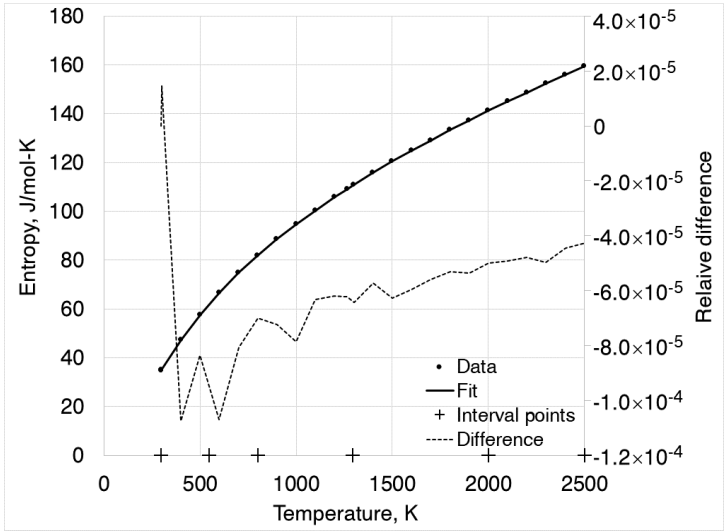


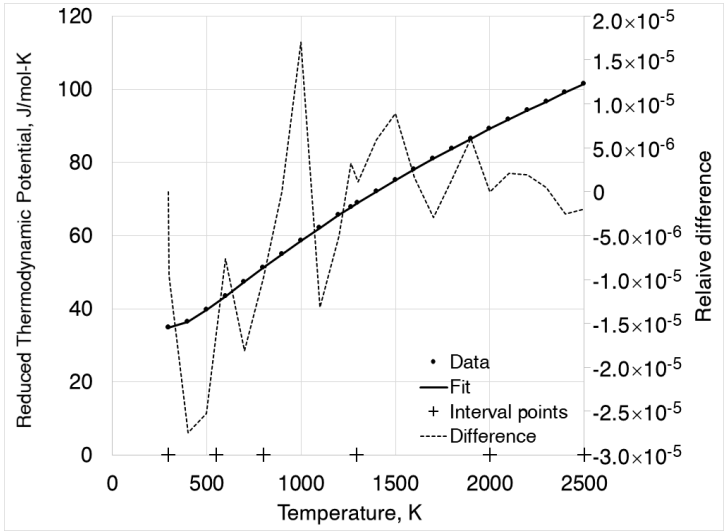
Titanium Oxide, Alpha (TiO) – O1Ti1(cr)_01

MW = 63.88 g/mol – $\Delta_f H^o$(298.15 K) = -542.665 kJ/mol – S^o(298.15 K) = 34.769 J/mol-K – p^o = 0.1 MPa



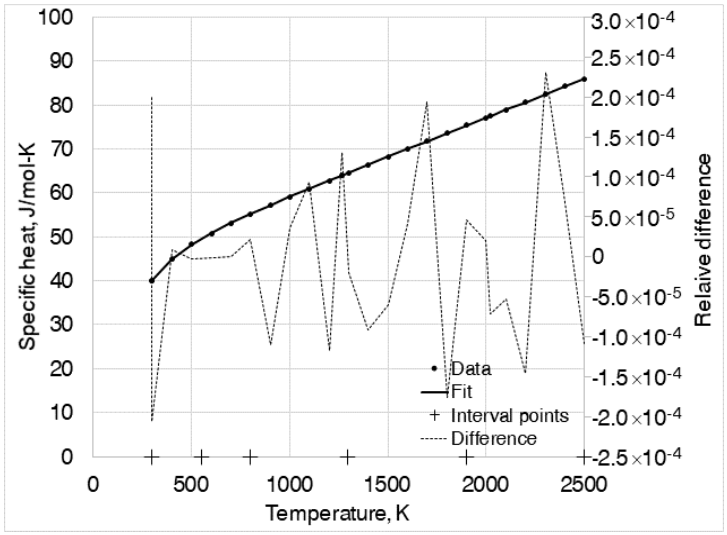


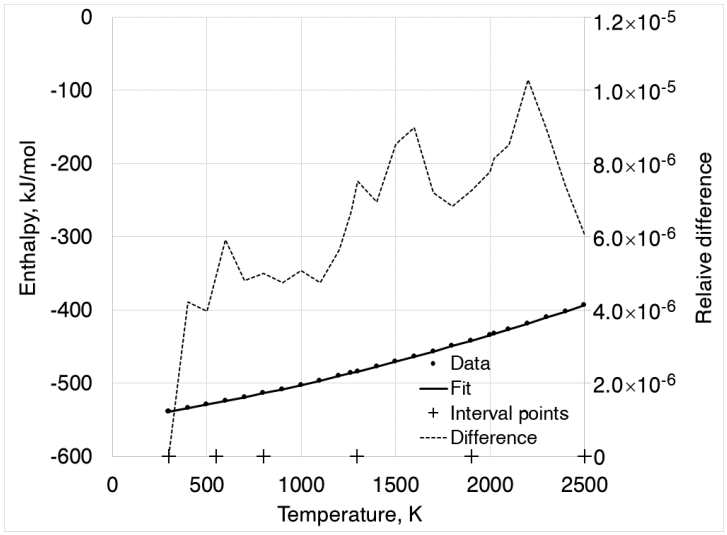


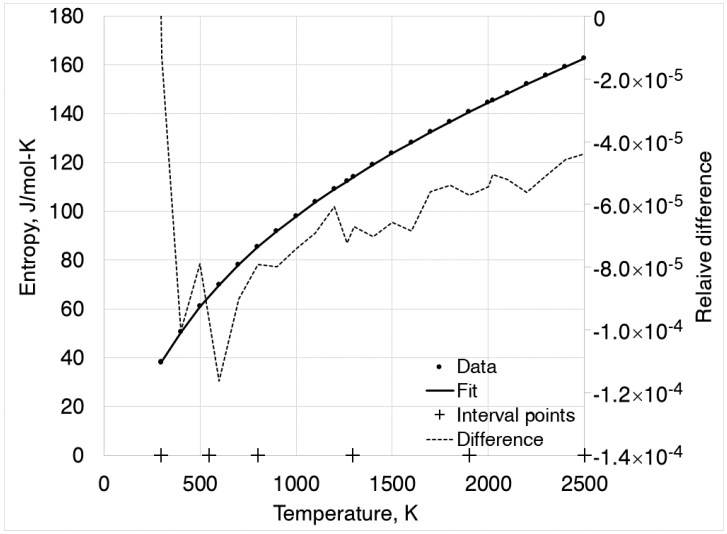


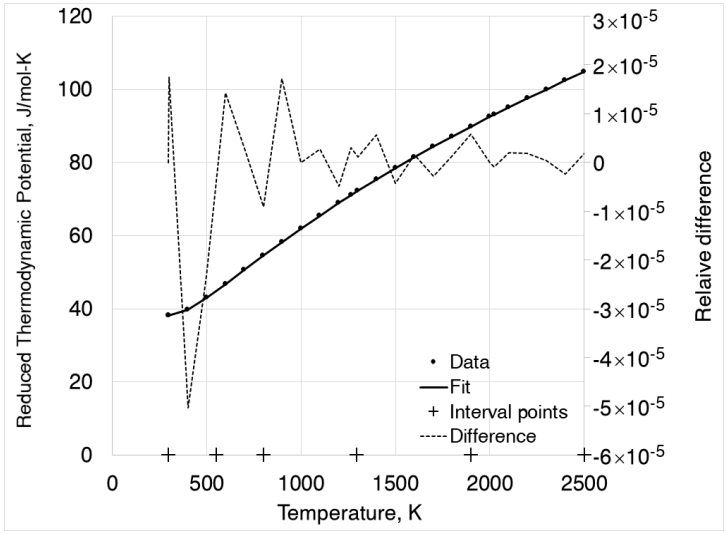
Titanium Oxide, Beta (TiO) – O1Ti1(cr)_02

MW = 63.88 g/mol – $\Delta_f H^o$(298.15 K) = -538.481 kJ/mol – S^o(298.15 K) = 38.077 J/mol-K – p^o = 0.1 MPa



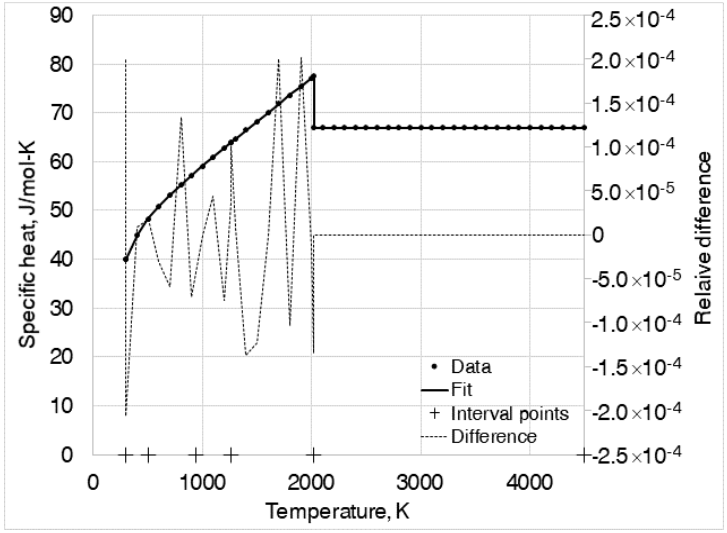


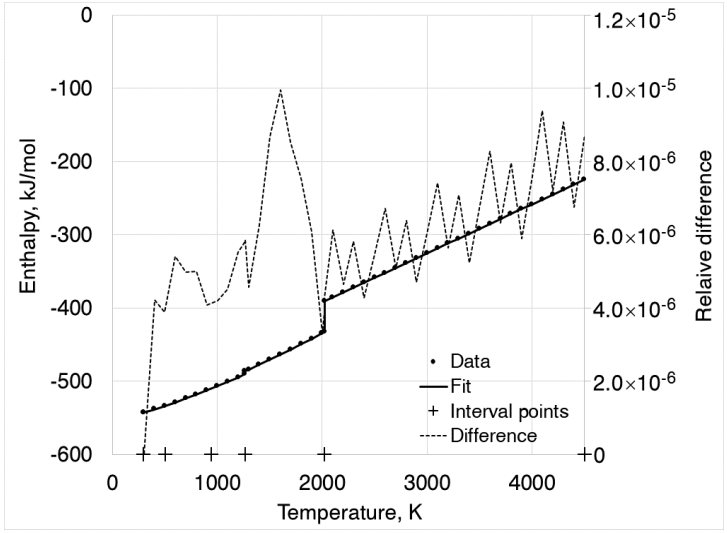


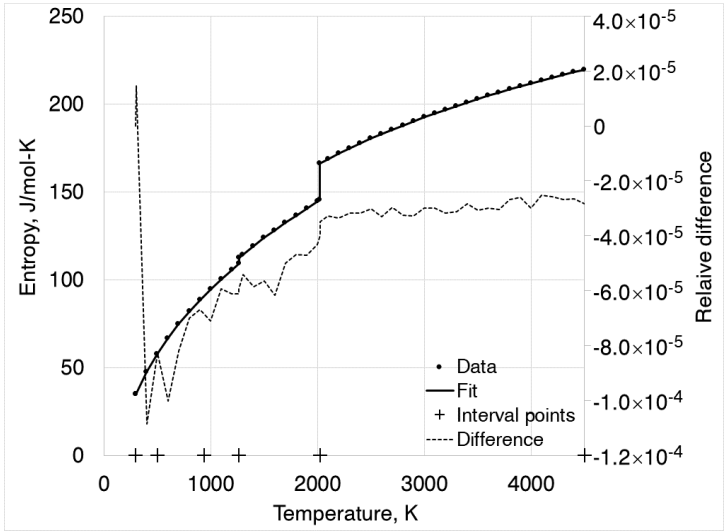


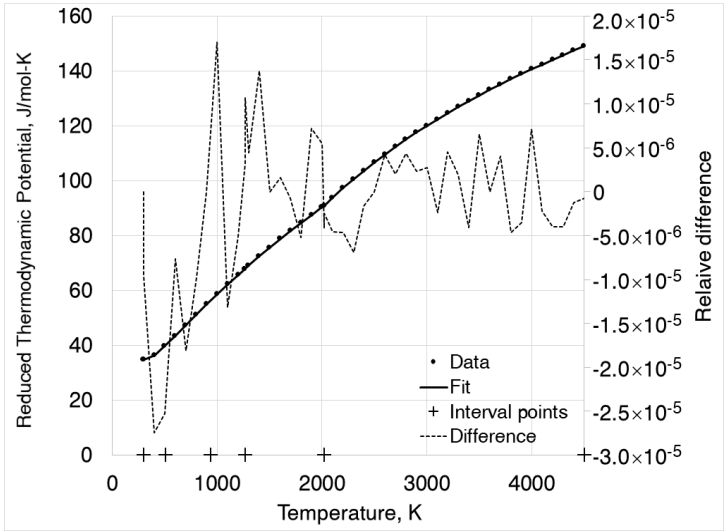
Titanium Oxide (TiO) – O1Ti1(cr,l)

MW = 63.88 g/mol – $\Delta_f H^o$(298.15 K) = -542.665 kJ/mol – S^o(298.15 K) = 34.769 J/mol-K – p^o = 0.1 MPa



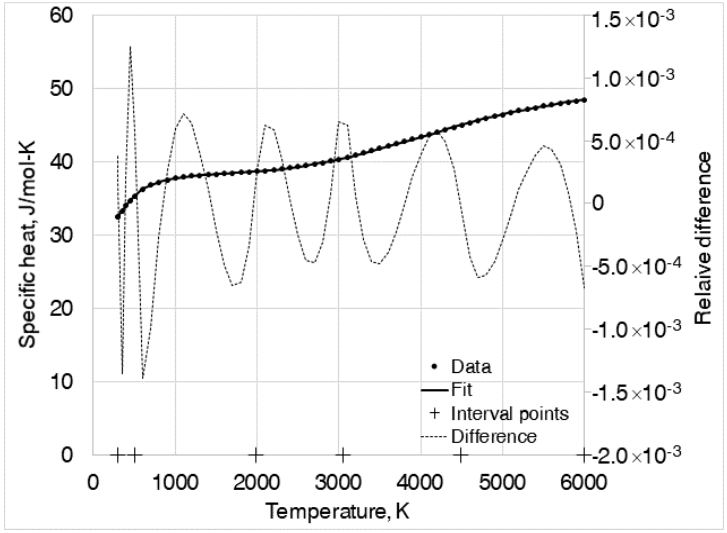


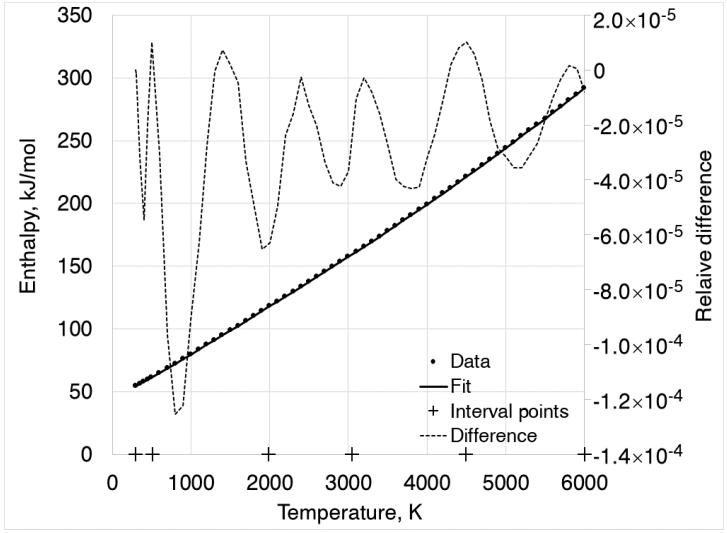


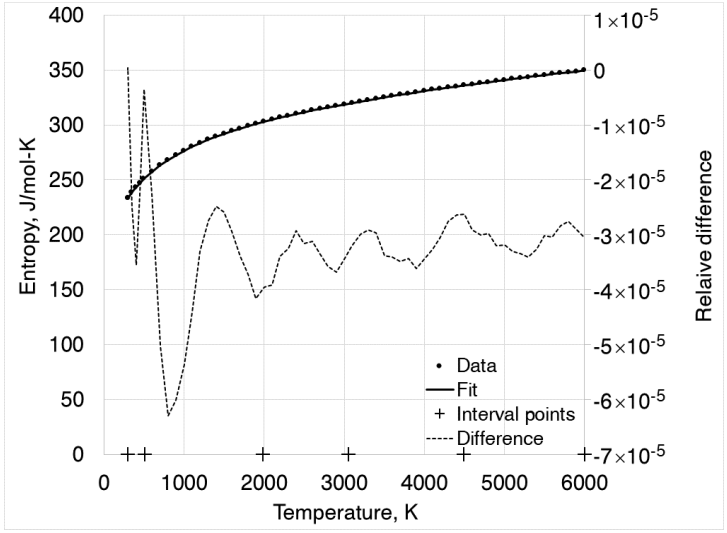


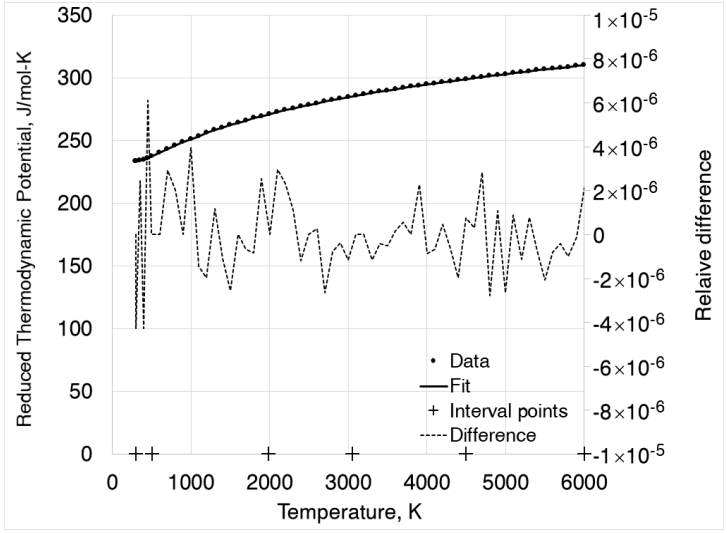
Titanium Oxide (TiO) – O1Ti1(g)

MW = 63.88 g/mol – $\Delta_f H^o$(298.15 K) = 54.392 kJ/mol – S^o(298.15 K) = 233.474 J/mol-K – p^o = 0.1 MPa



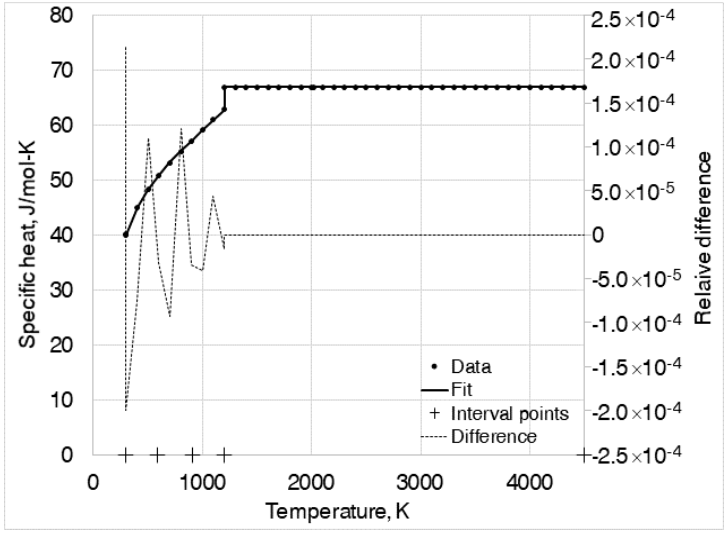


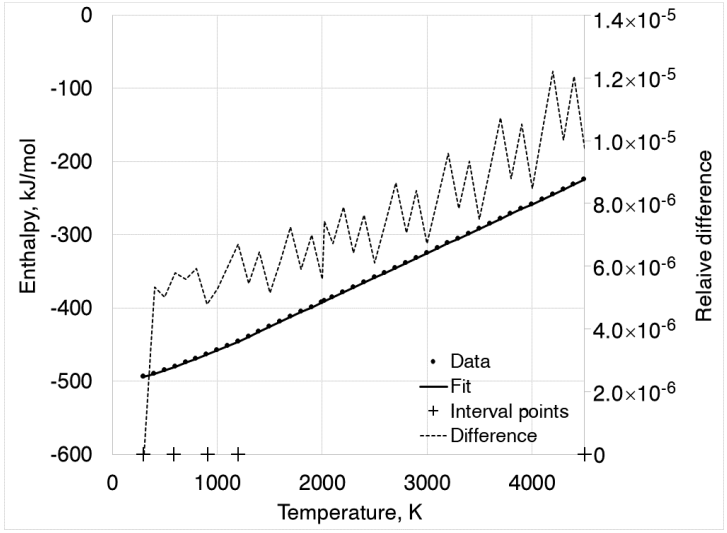


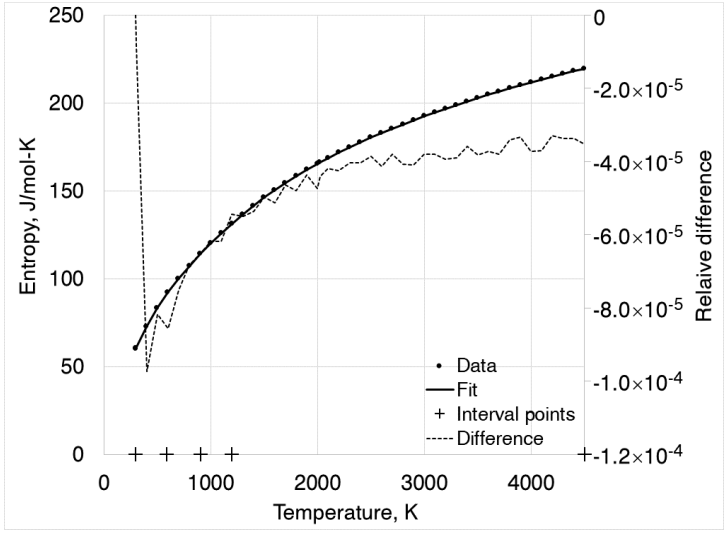


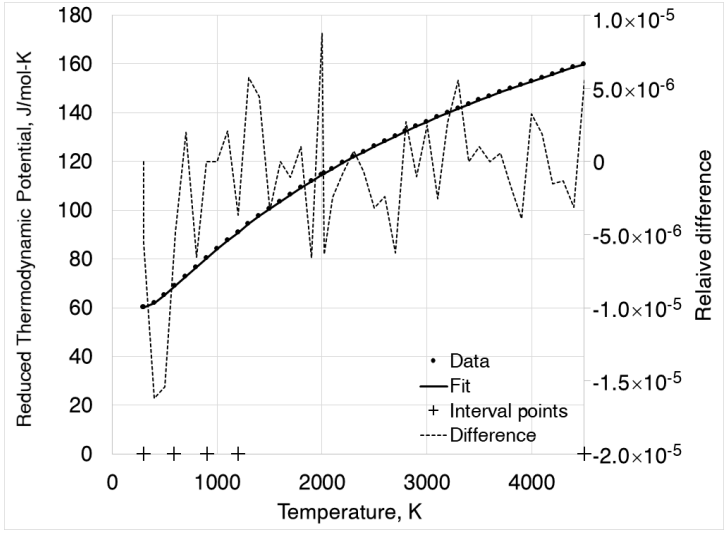
Titanium Oxide (TiO) – O1Ti1(l)

MW = 63.88 g/mol – $\Delta_f H^o$(298.15 K) = -493.939 kJ/mol – S^o(298.15 K) = 60.14 J/mol-K – p^o = 0.1 MPa



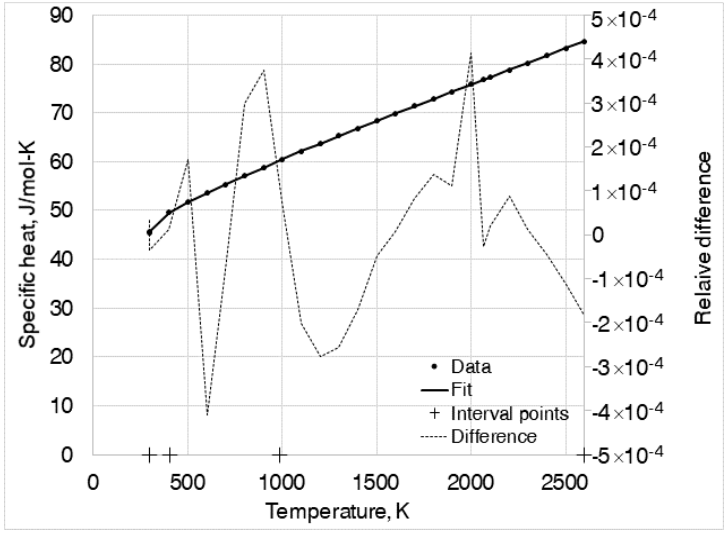


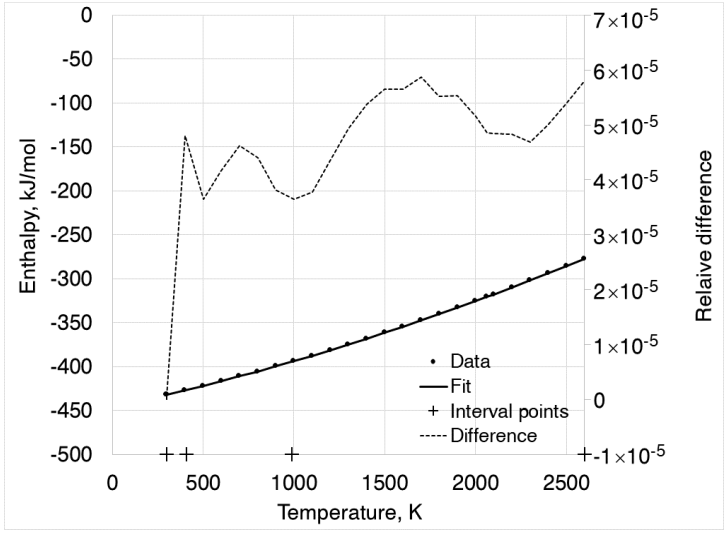


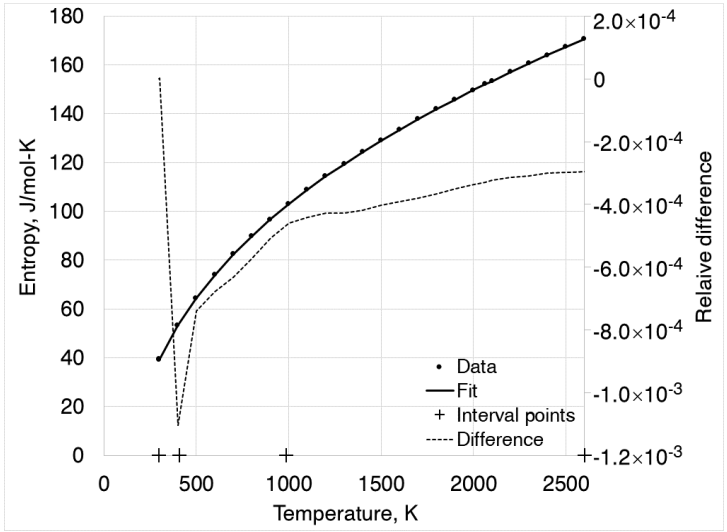


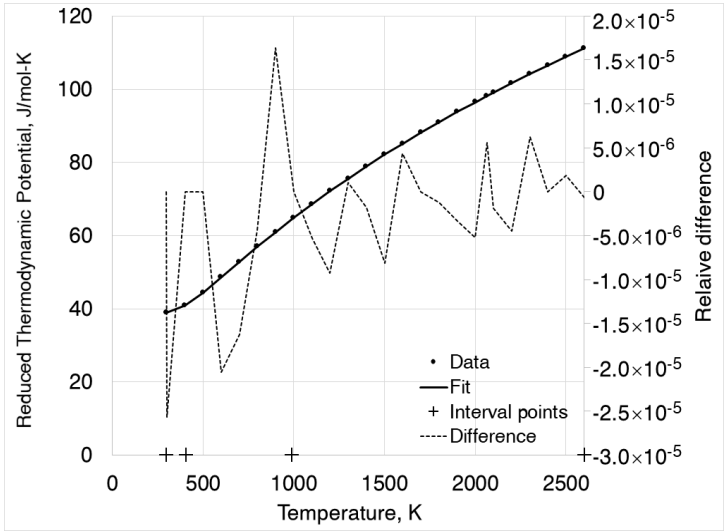
Vanadium Oxide (VO) – O1V1(cr)

MW = 66.9409 g/mol – $\Delta_f H^o$(298.15 K) = -431.789 kJ/mol – S^o(298.15 K) = 39.011 J/mol-K – p^o = 0.1 MPa



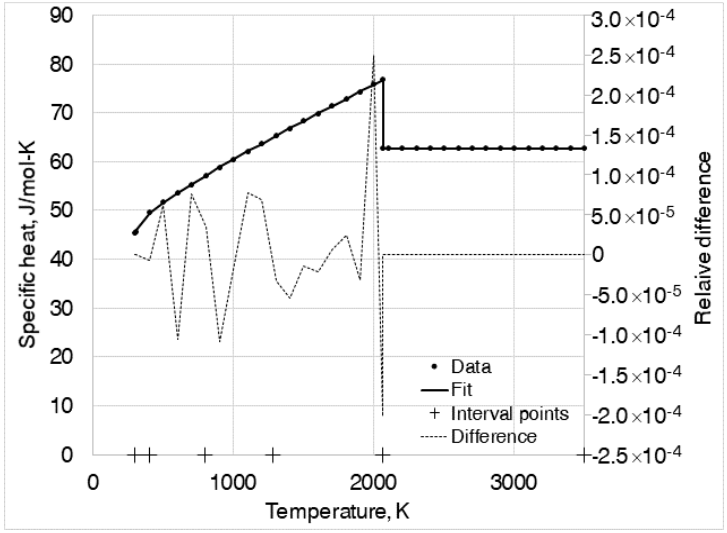


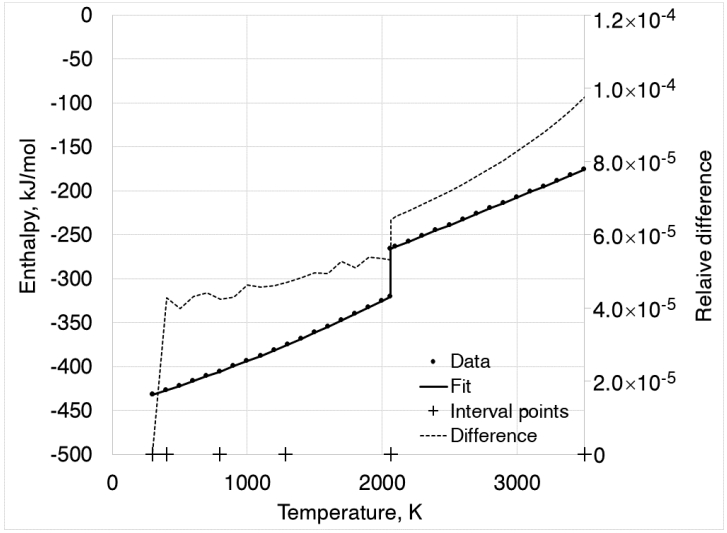


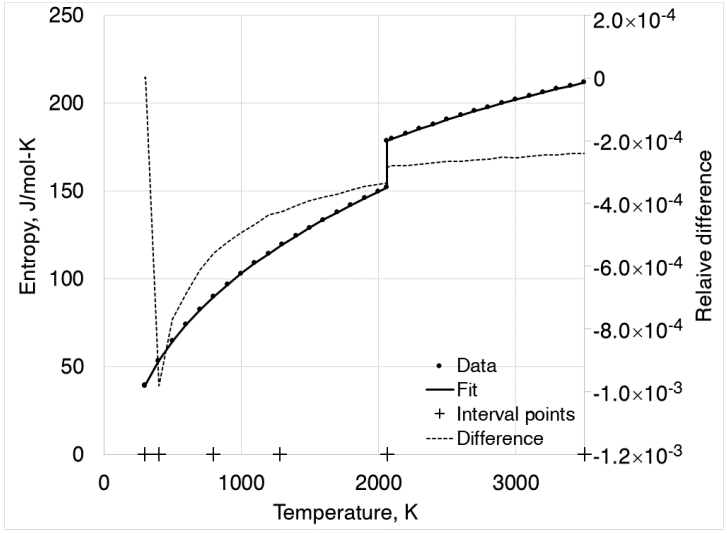


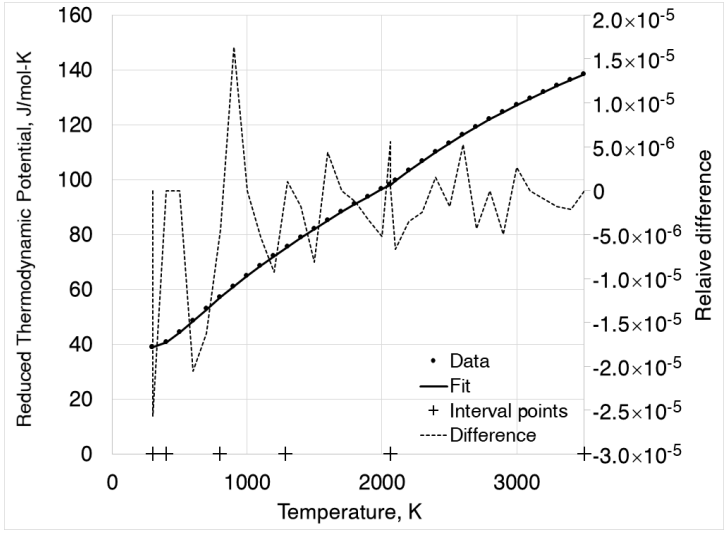
Vanadium Oxide (VO) – O1V1(cr,l)

MW = 66.9409 g/mol – $\Delta_f H^o$(298.15 K) = -431.789 kJ/mol – S^o(298.15 K) = 39.011 J/mol-K – p^o = 0.1 MPa



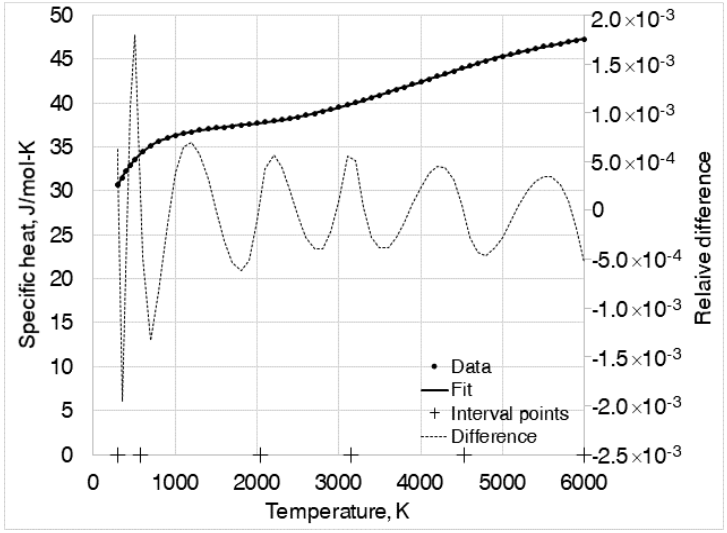


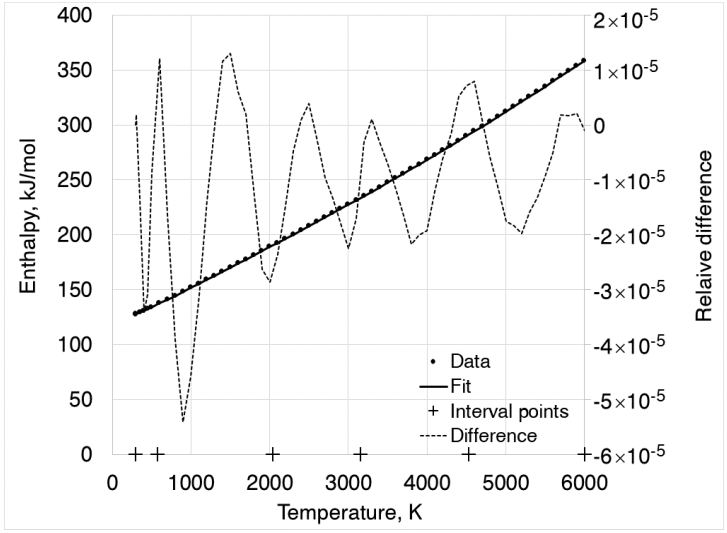


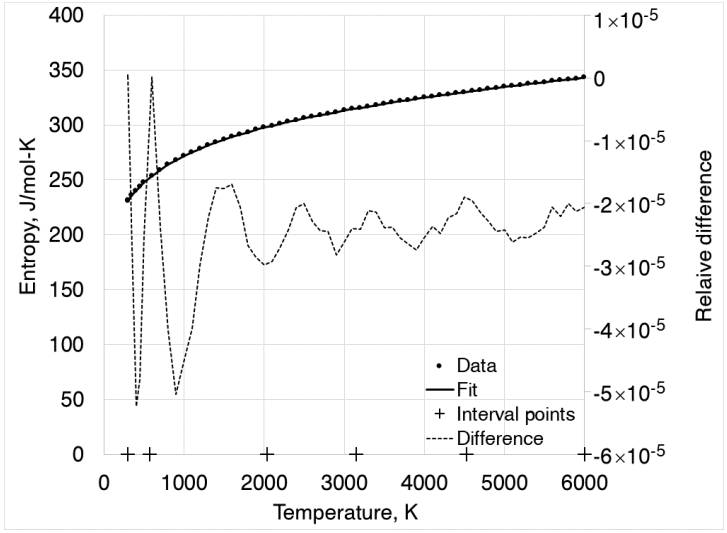


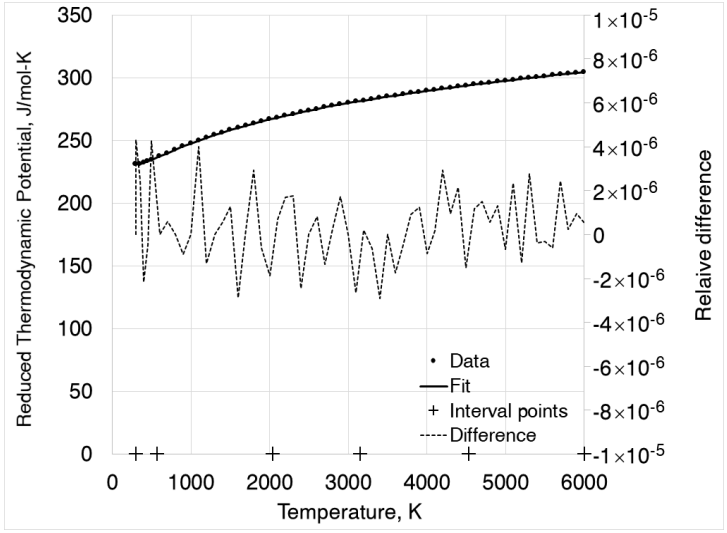
Vanadium Oxide (VO) – O1V1(g)

MW = 66.9409 g/mol – $\Delta_f H^o$(298.15 K) = 127.612 kJ/mol – S^o(298.15 K) = 230.901 J/mol-K – p^o = 0.1 MPa



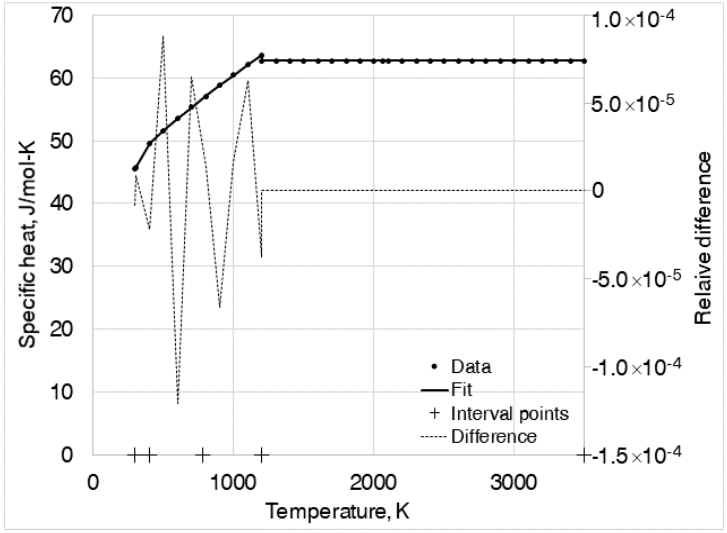


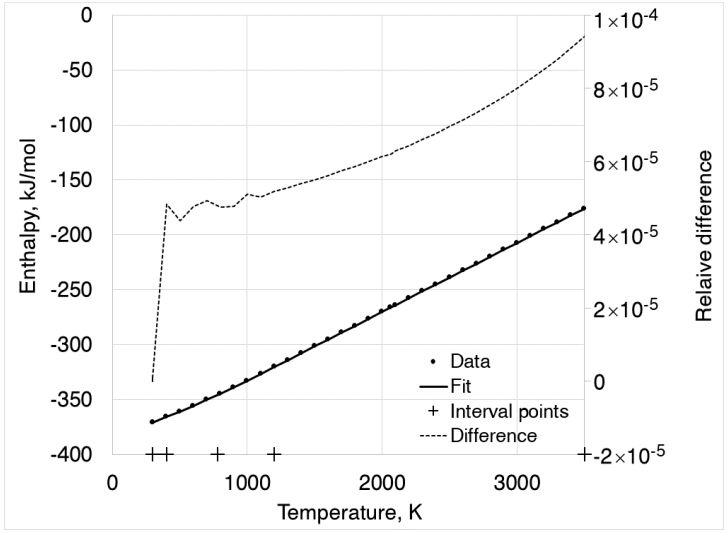


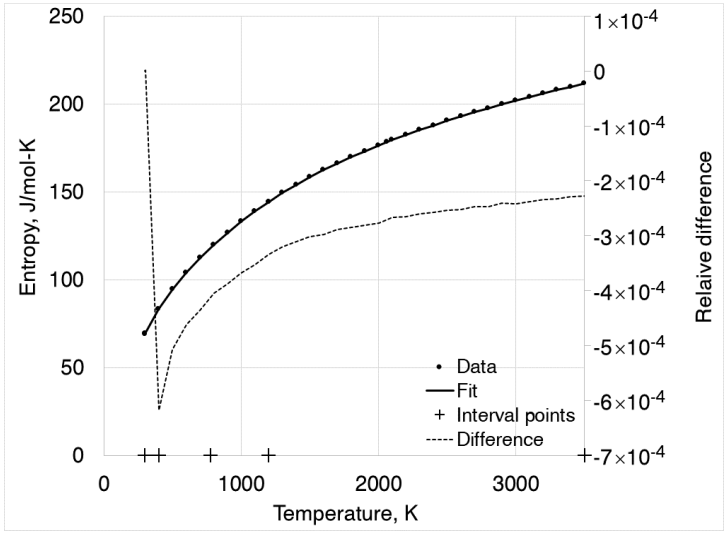


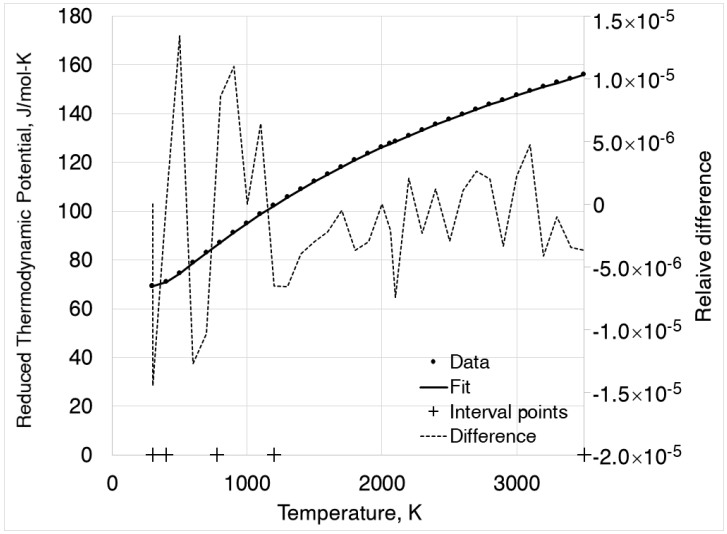
Vanadium Oxide (VO) – O1V1(l)

MW = 66.9409 g/mol – $\Delta_f H^o$(298.15 K) = -370.901 kJ/mol – S^o(298.15 K) = 69.137 J/mol-K – p^o = 0.1 MPa



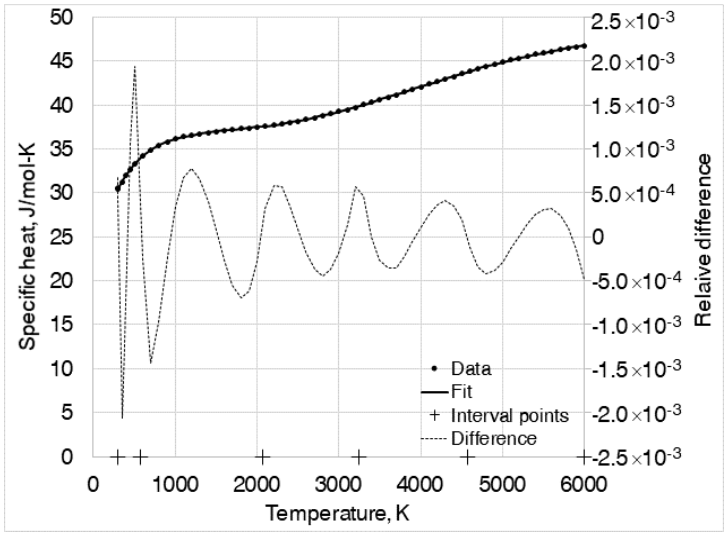


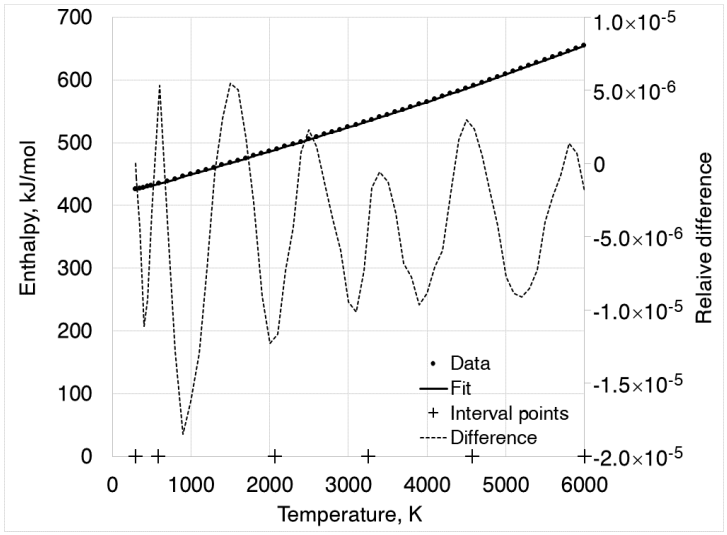


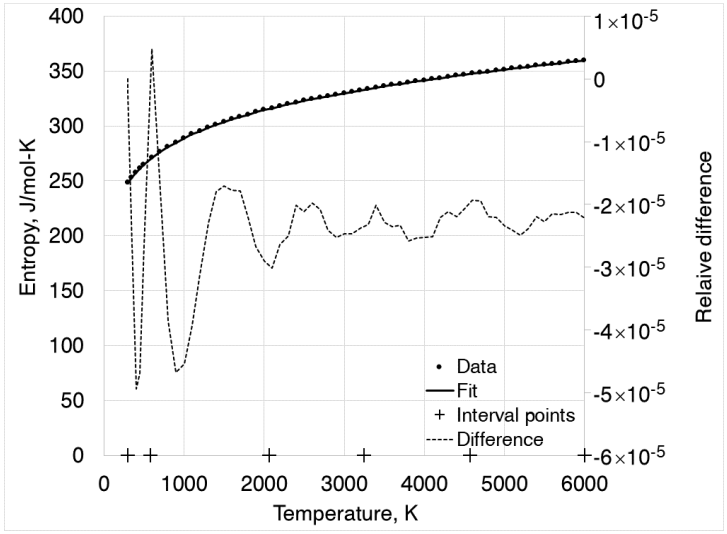


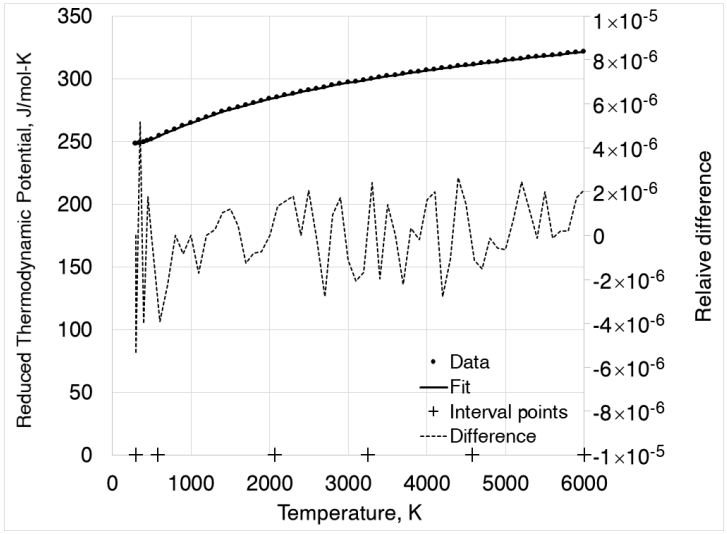
Tungsten Oxide (WO) – O1W1(g)

MW = 199.85 g/mol – $\Delta_f H^o$(298.15 K) = 425.094 kJ/mol – S^o(298.15 K) = 248.198 J/mol-K – p^o = 0.1 MPa



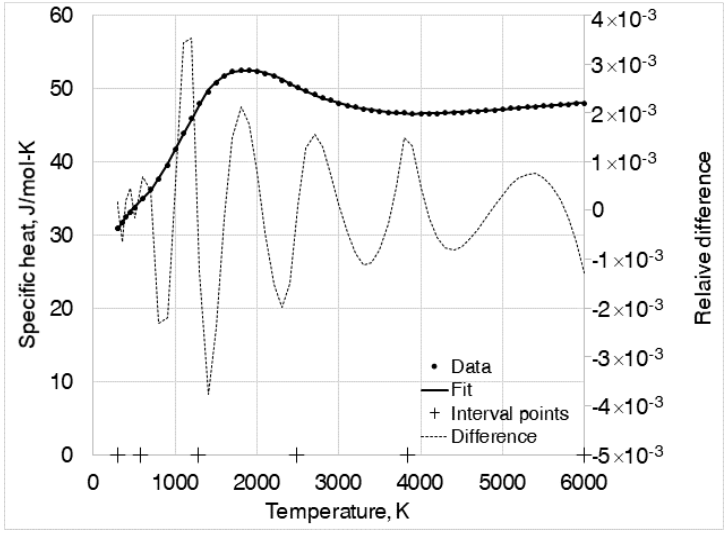


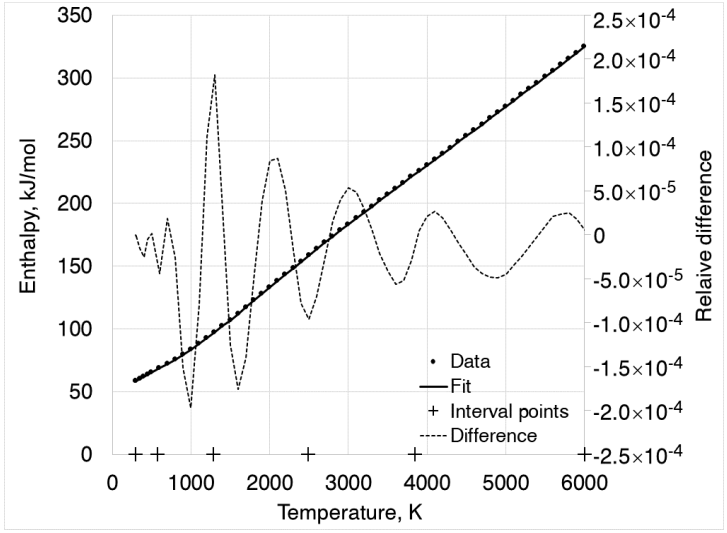


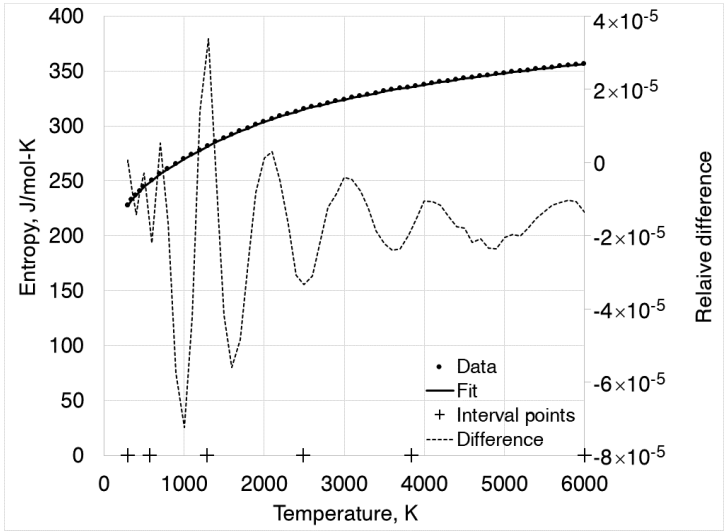


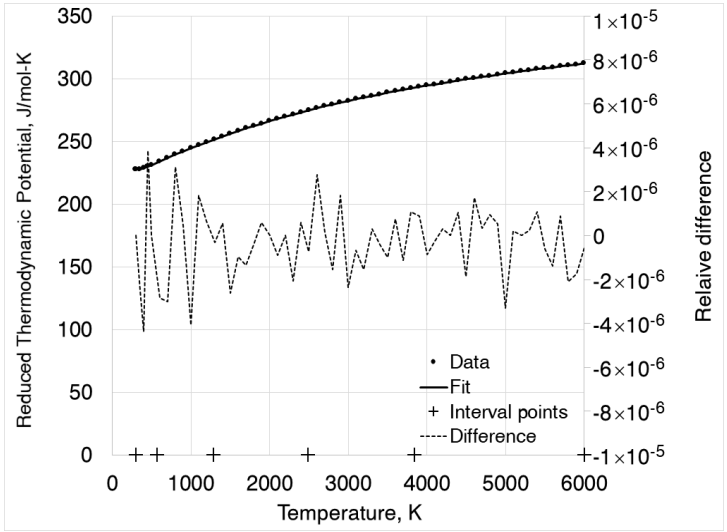
Zirconium Oxide (ZrO) – O1Zr1(g)

MW = 107.22 g/mol – $\Delta_f H^o$(298.15 K) = 58.576 kJ/mol – S^o(298.15 K) = 227.627 J/mol-K – p^o = 0.1 MPa



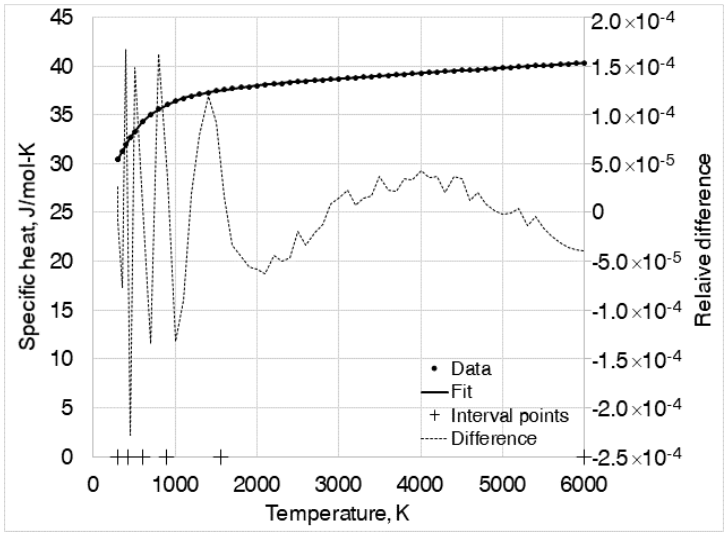


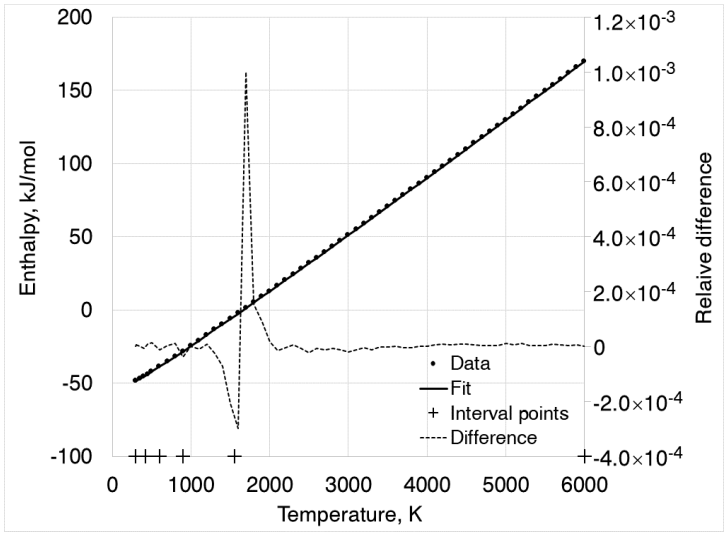


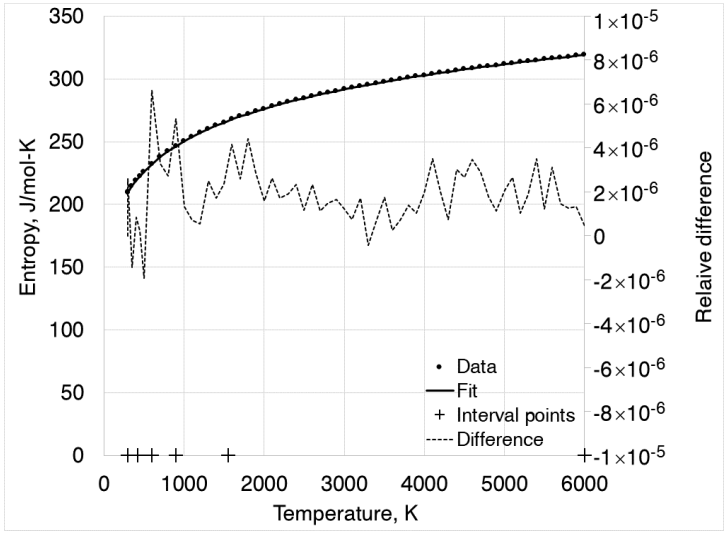


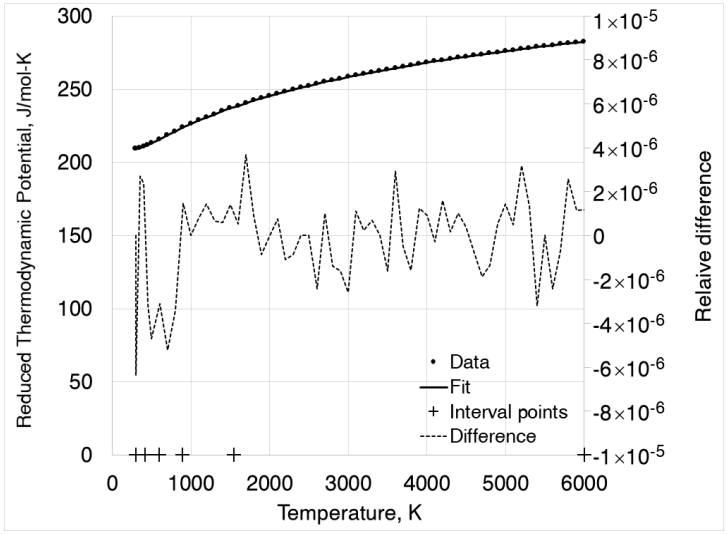
Oxygen, Ion (O_2^-) – O2-(g)

MW = 31.9993 g/mol – $\Delta_f H^o$(298.15 K) = -48.593 kJ/mol – S^o(298.15 K) = 209.591 J/mol-K – p^o = 0.1 MPa



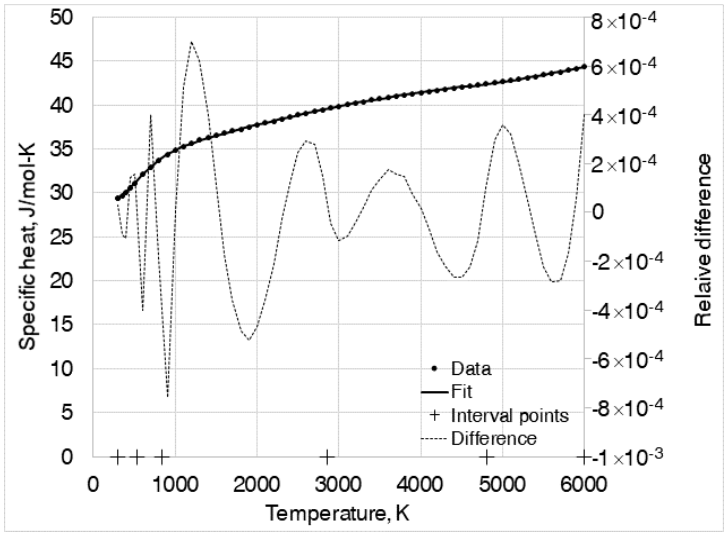


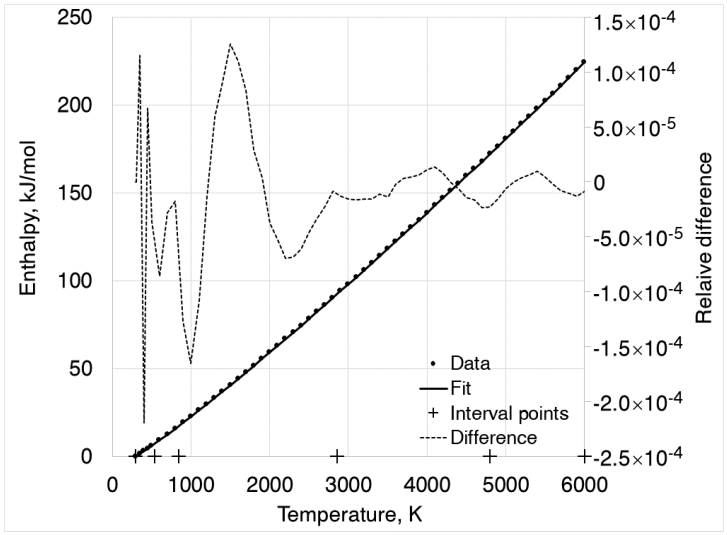


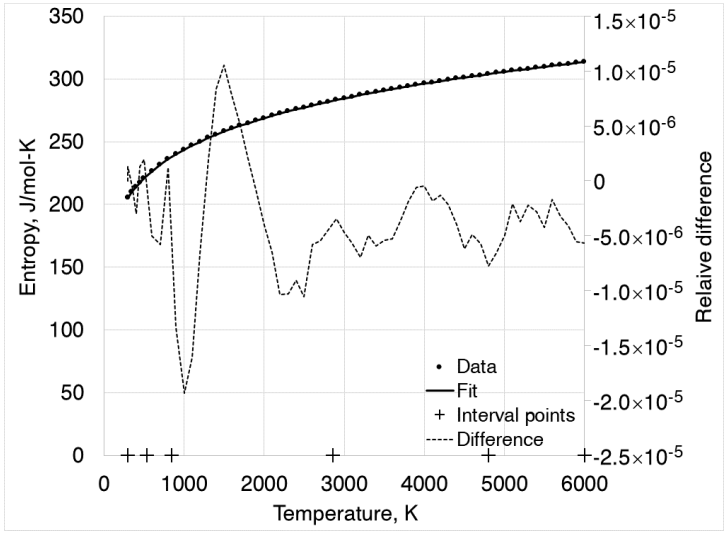


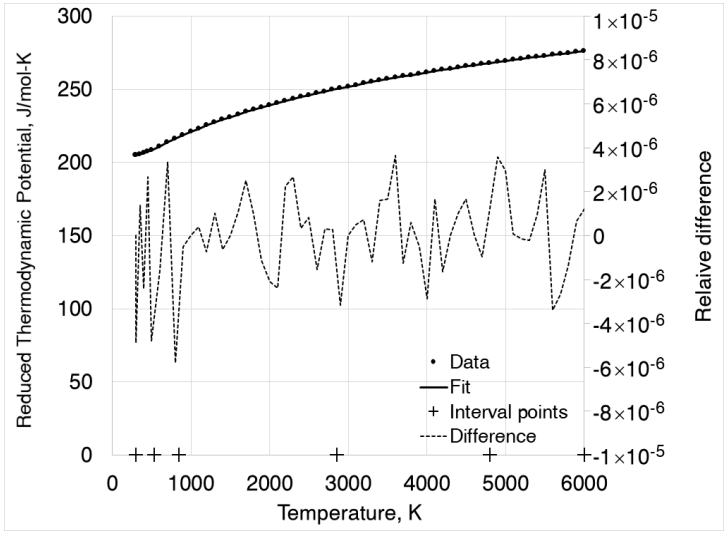
Oxygen (O_2) – O2(ref)

MW = 31.9988 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 205.147 J/mol-K – p^o = 0.1 MPa



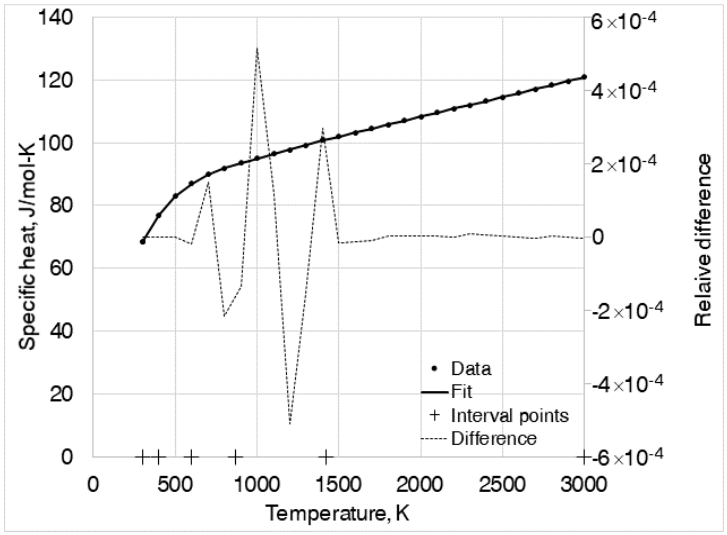


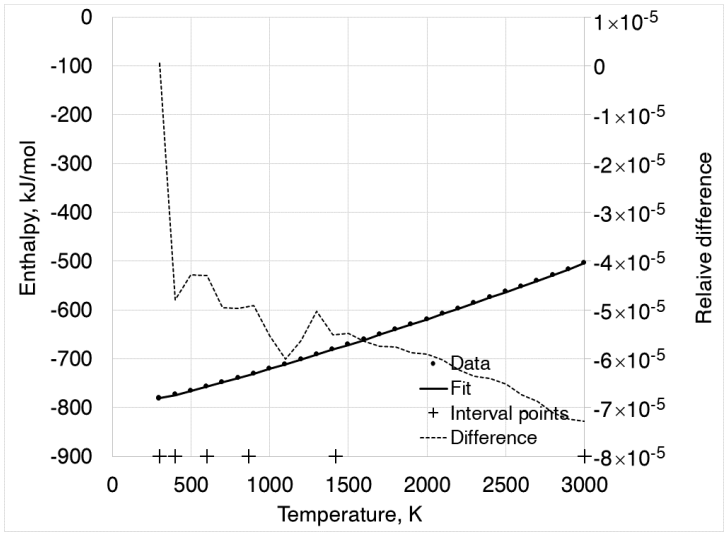


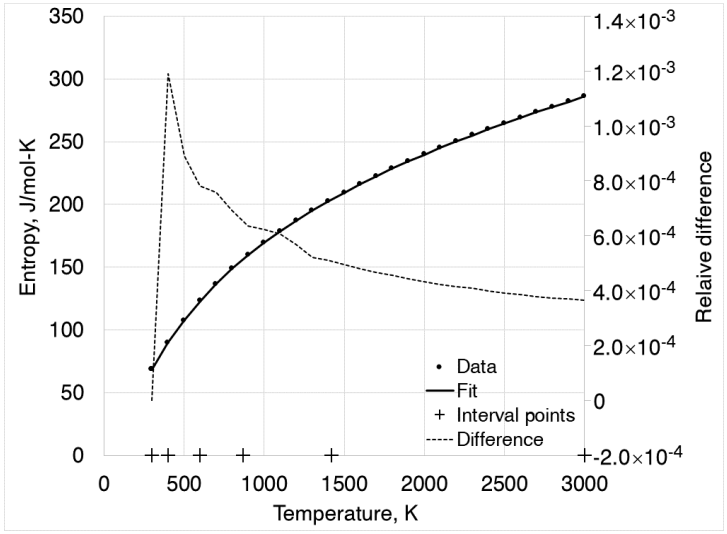


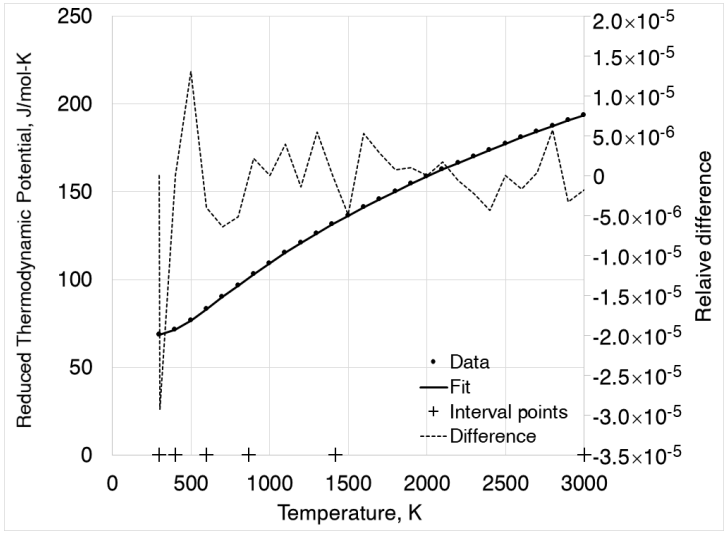
Tungsten Oxide ($WO_{2.72}$) – O2.72W1(cr)

MW = 227.37 g/mol – $\Delta_f H^o$(298.15 K) = -781.153 kJ/mol – S^o(298.15 K) = 68.425 J/mol-K – p^o = 0.1 MPa



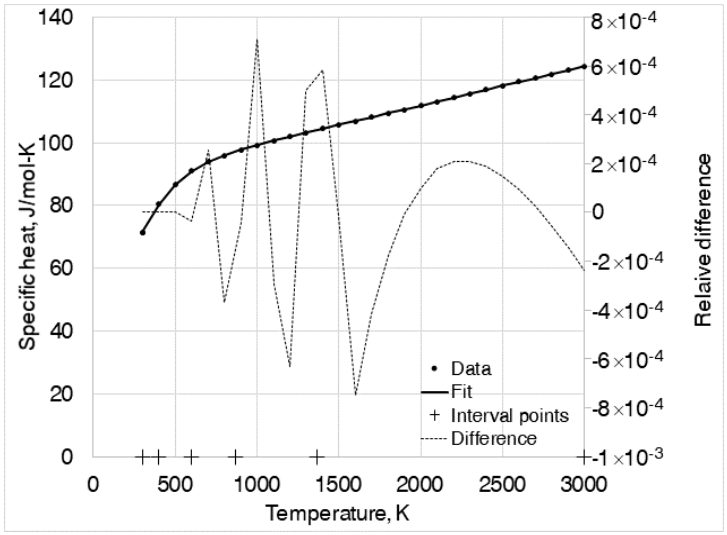


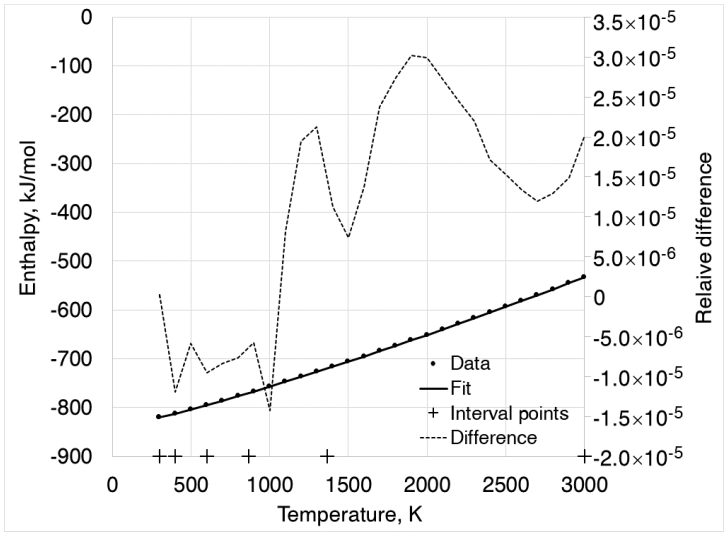


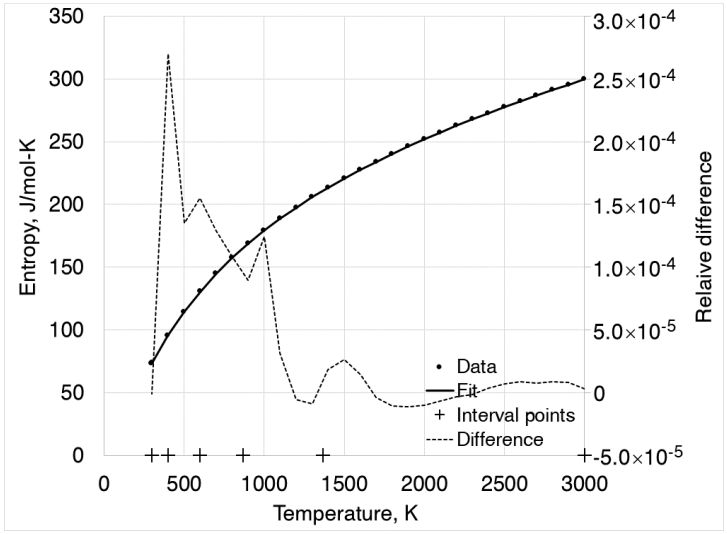


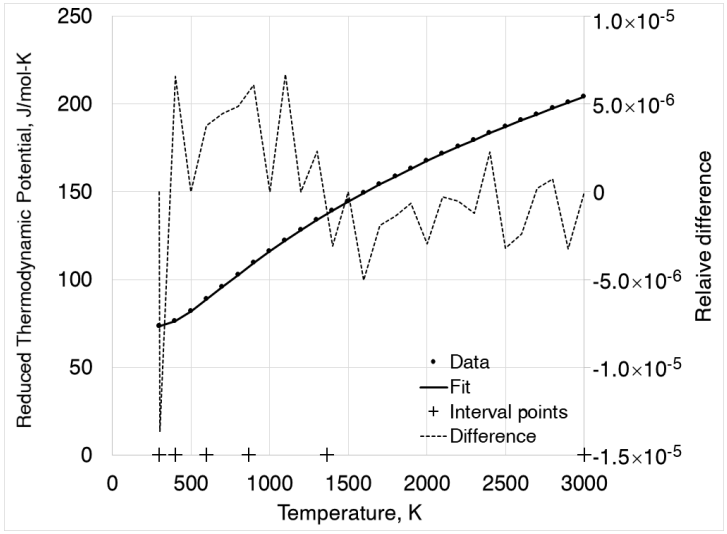
Tungsten Oxide ($WO_{2.90}$) – O2.90W1(cr)

MW = 230.25 g/mol – $\Delta_f H^o$(298.15 K) = -820.064 kJ/mol – S^o(298.15 K) = 73.371 J/mol-K – p^o = 0.1 MPa



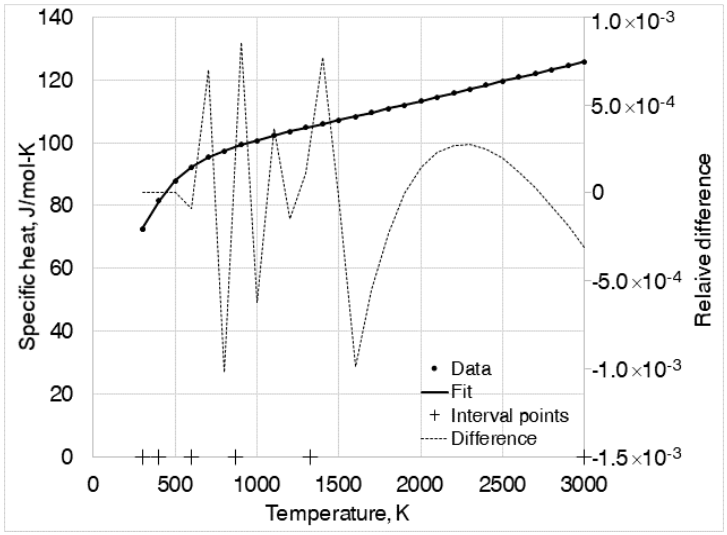


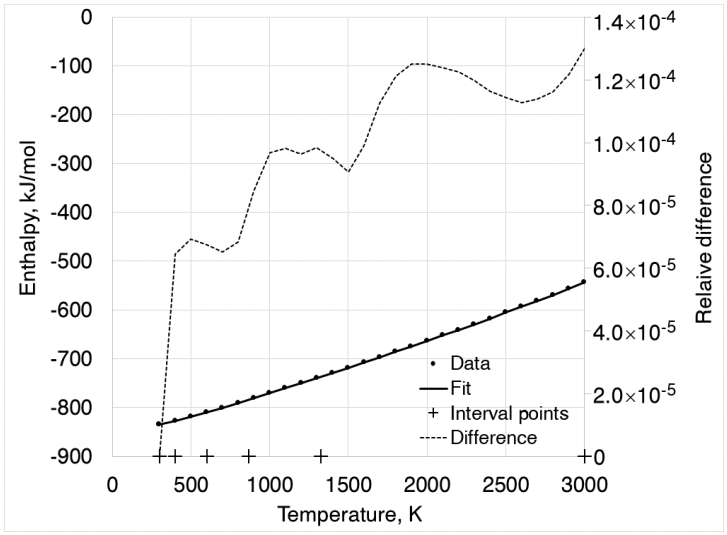


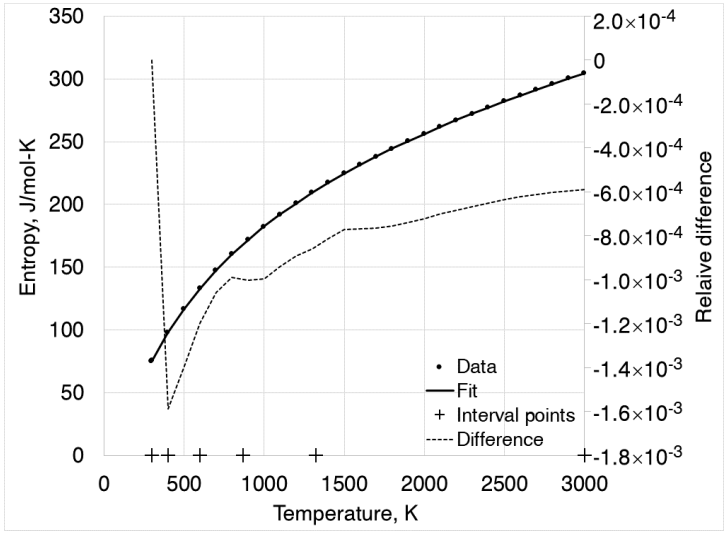


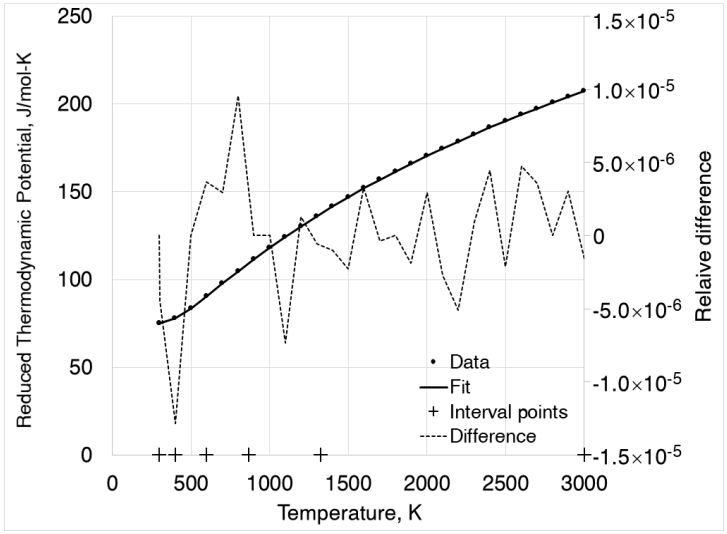
Tungsten Oxide ($WO_{2.96}$) – O2.96W1(cr)

MW = 231.21 g/mol – $\Delta_f H^o$(298.15 K) = -834.959 kJ/mol – S^o(298.15 K) = 74.894 J/mol-K – p^o = 0.1 MPa



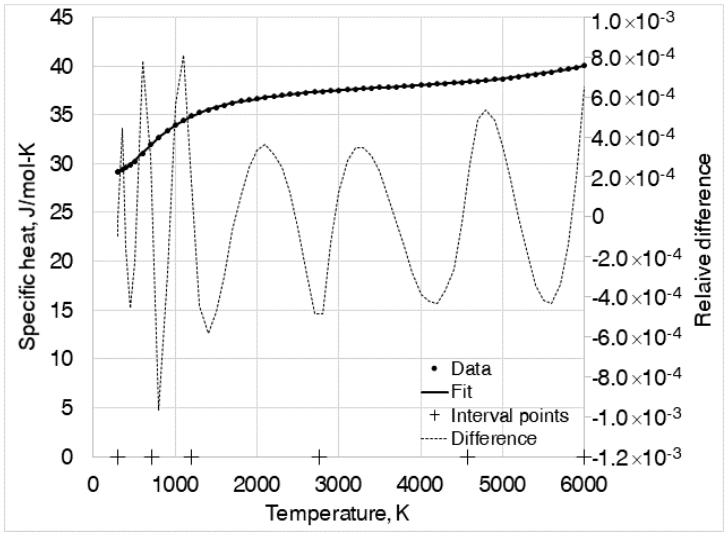


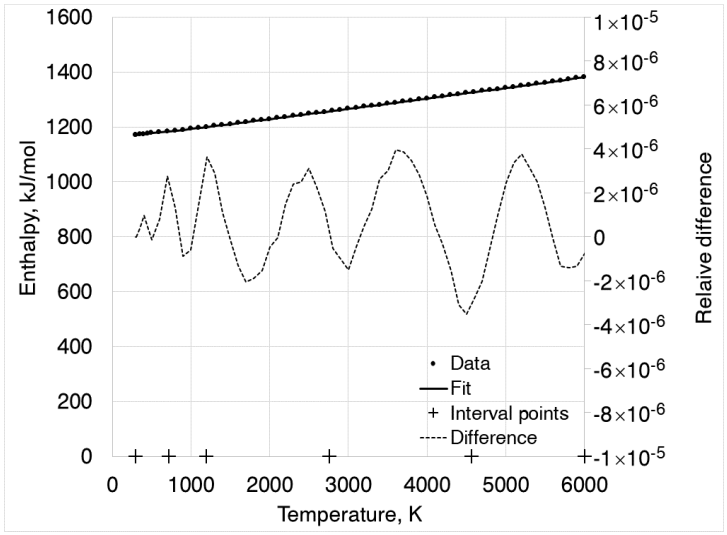


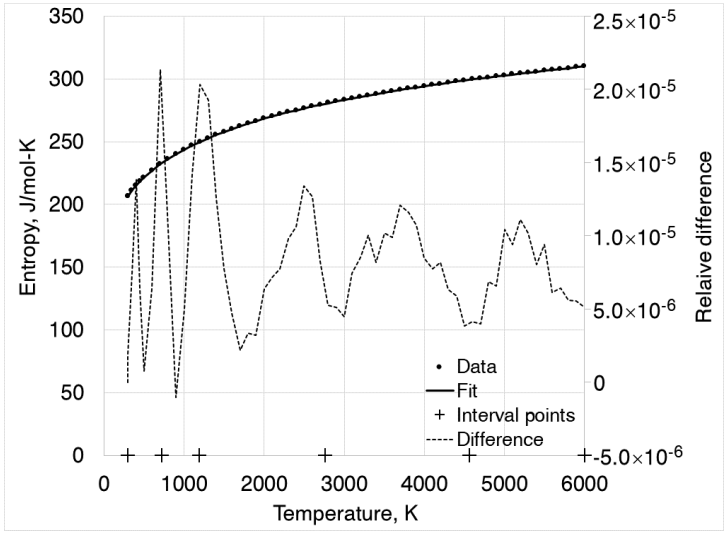


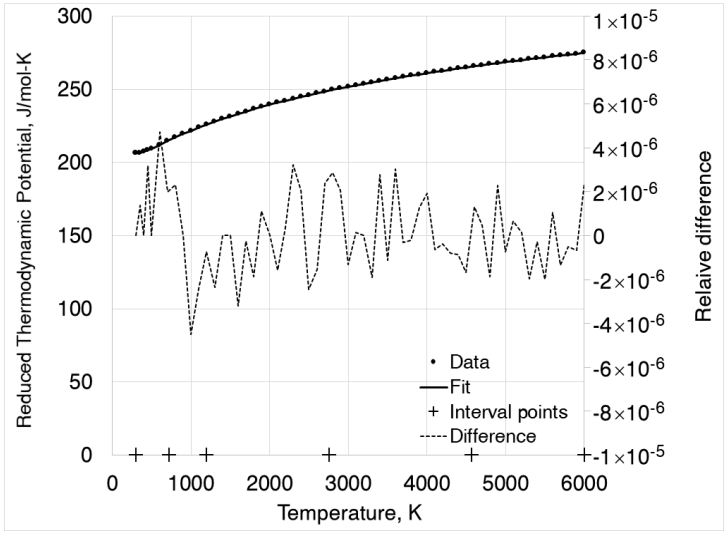
Oxygen, Ion (O_2^+) – O2+(g)

MW = 31.9983 g/mol – $\Delta_f H^o$(298.15 K) = 1170.888 kJ/mol – S^o(298.15 K) = 206.215 J/mol-K – p^o = 0.1 MPa



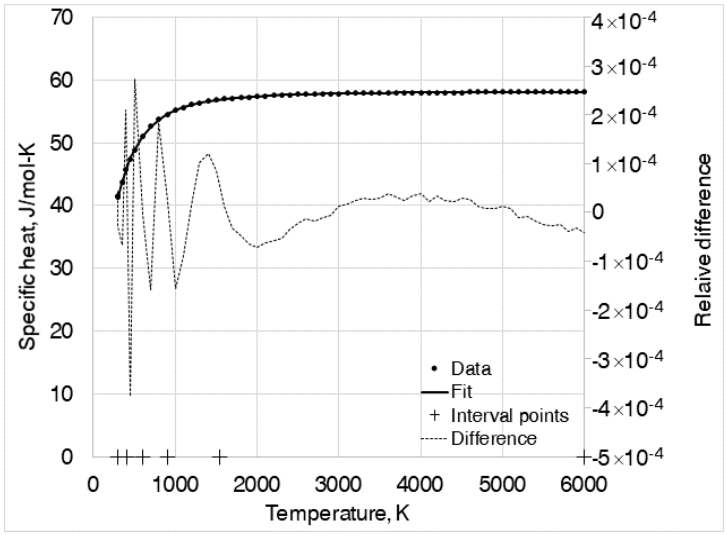


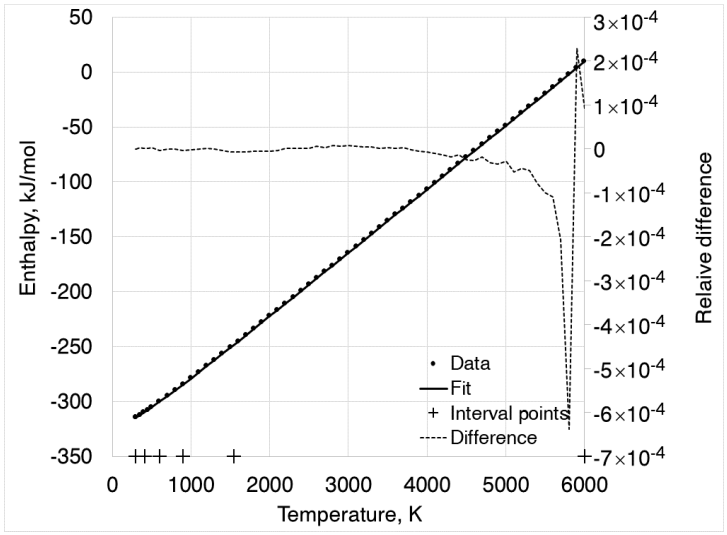


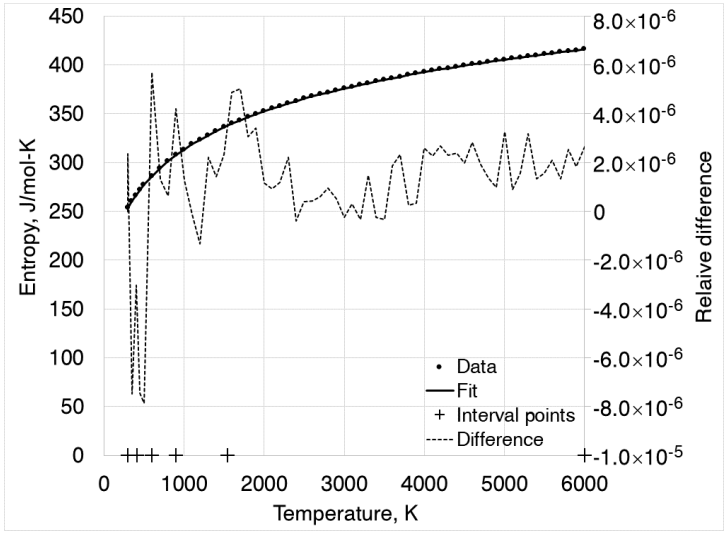


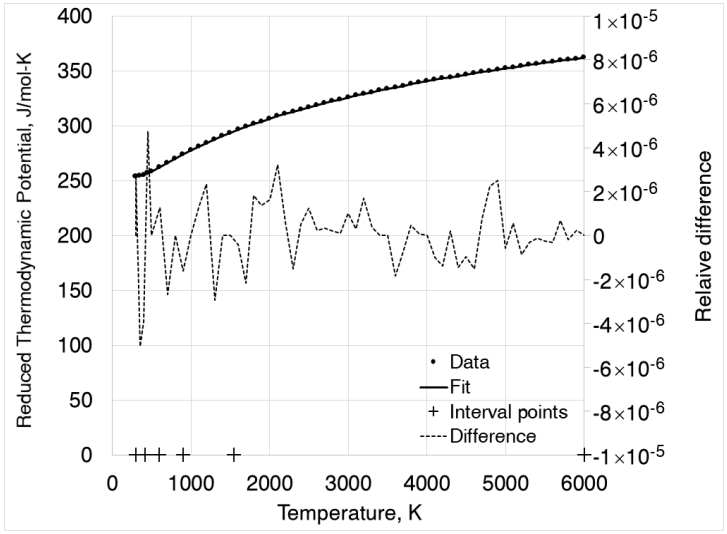
Phosphorus Oxide (PO_2) – O2P1(g)

MW = 62.9726 g/mol – $\Delta_f H^o$(298.15 K) = -314.524 kJ/mol – S^o(298.15 K) = 253.685 J/mol-K – p^o = 0.1 MPa



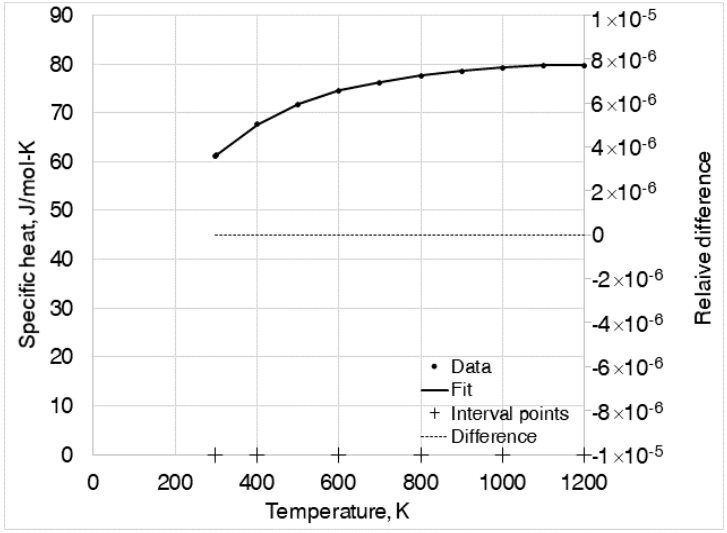


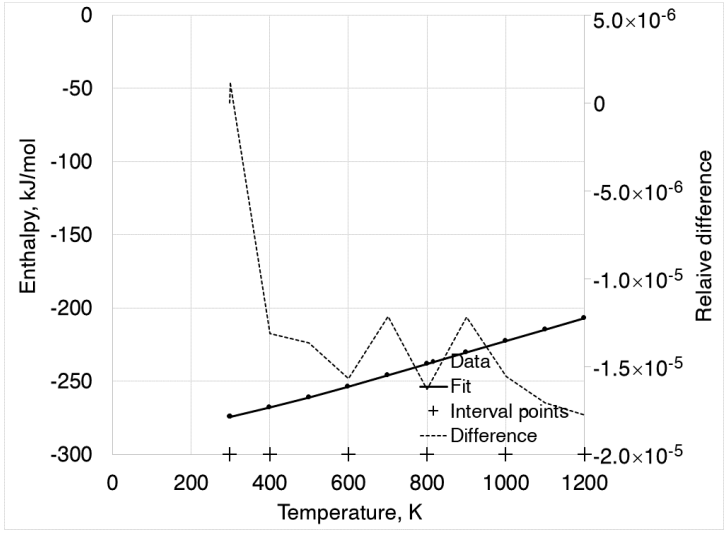


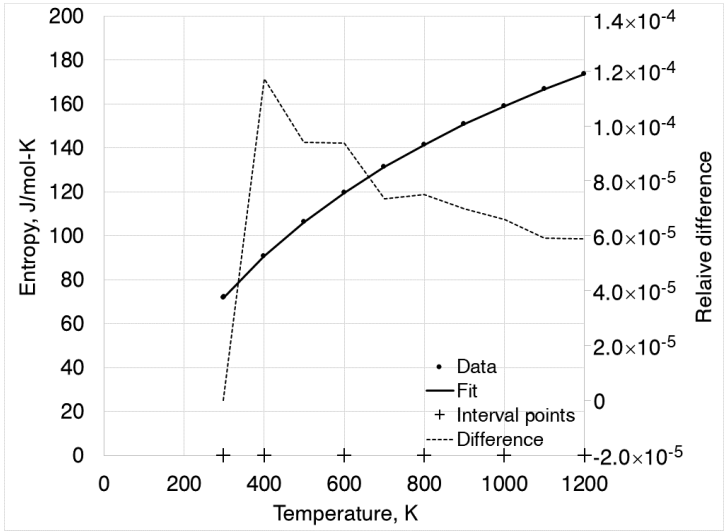


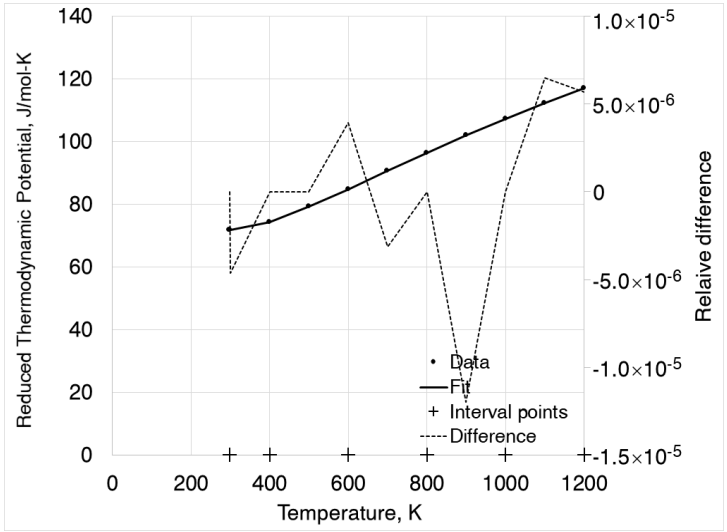
Lead Oxide (PbO_2) – O2Pb1(cr)

MW = 239.2 g/mol – $\Delta_f H^o$(298.15 K) = -274.47 kJ/mol – S^o(298.15 K) = 71.797 J/mol-K – p^o = 0.1 MPa



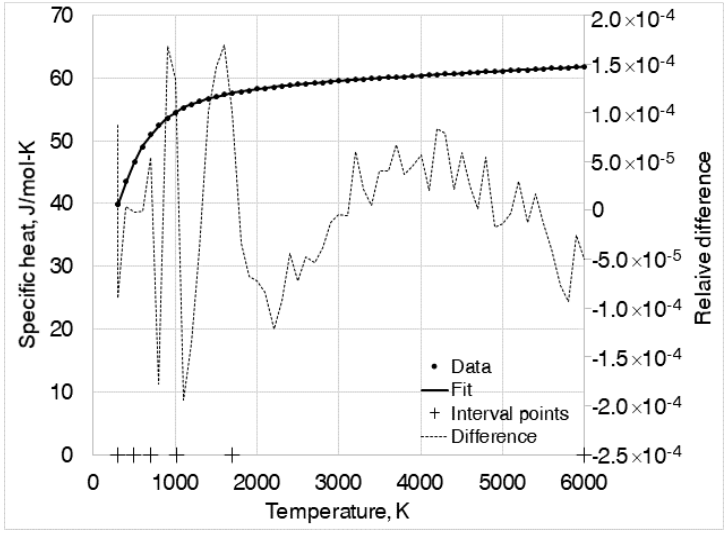


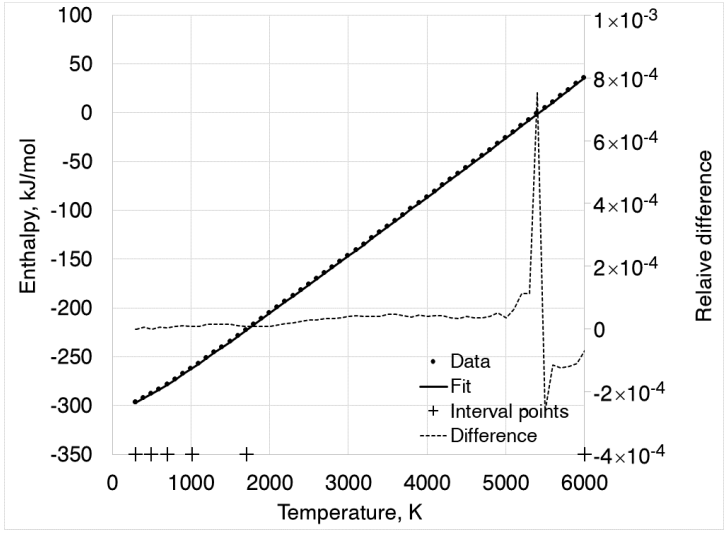


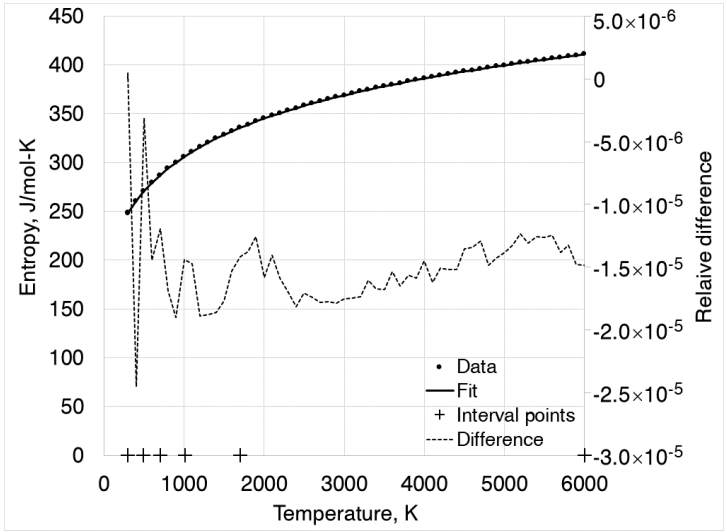


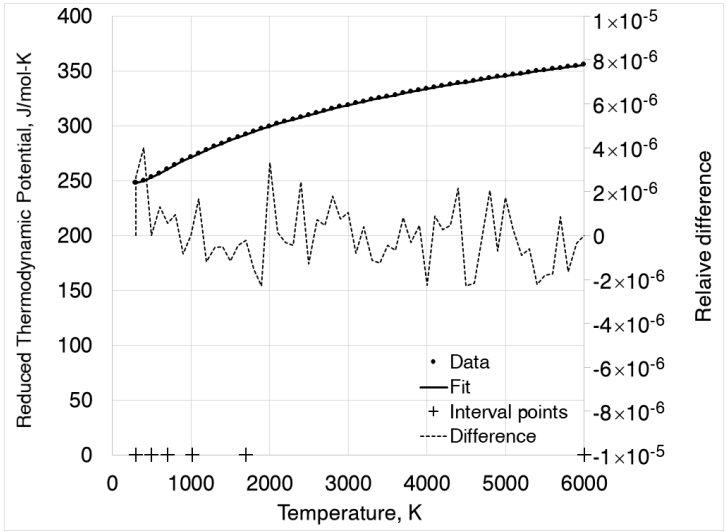
Sulfur Dioxide (SO_2) – O2S1(g)

MW = 64.06 g/mol – $\Delta_f H^o$(298.15 K) = -296.842 kJ/mol – S^o(298.15 K) = 248.212 J/mol-K – p^o = 0.1 MPa



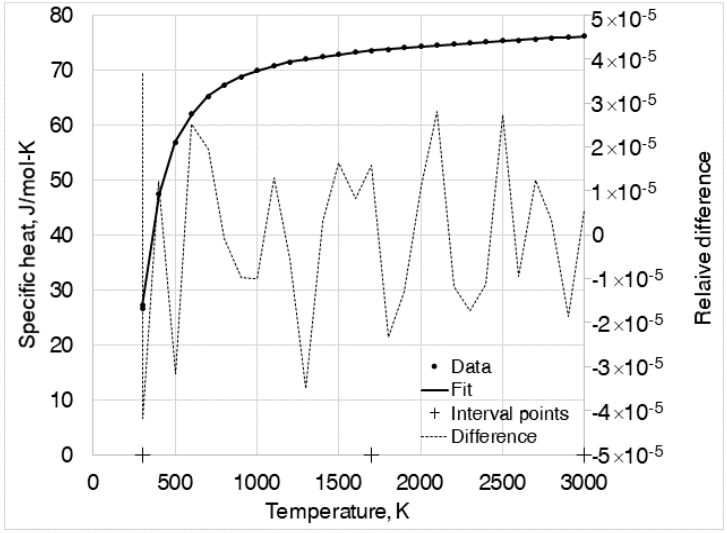


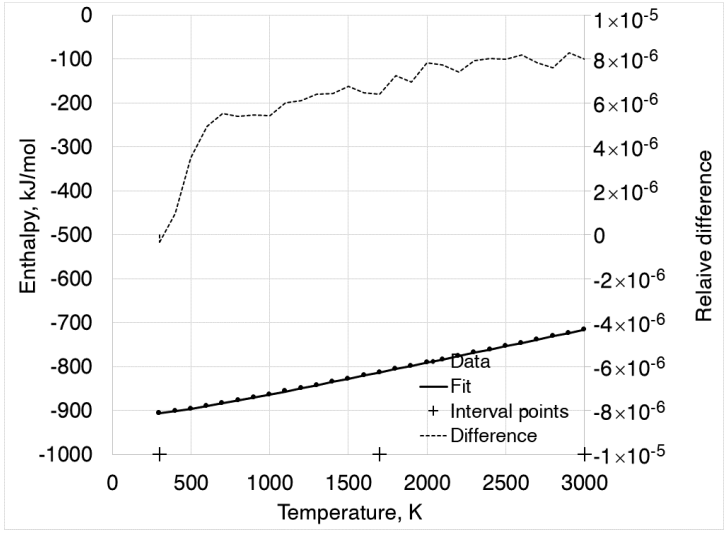


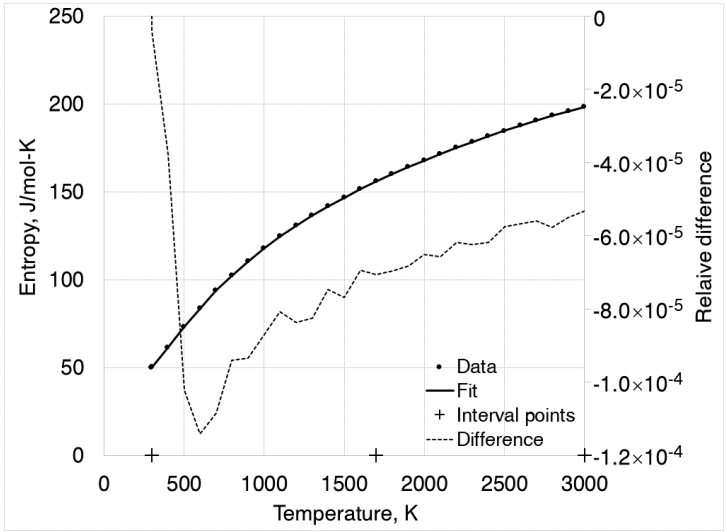


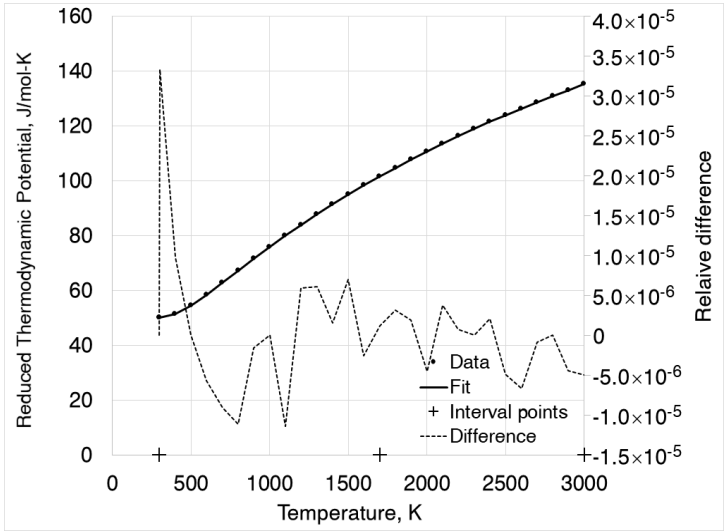
Silicon Oxide, Cristobalite, High (SiO_2) – O2Si1(cr)_01

MW = 60.0843 g/mol – $\Delta_f H^o$(298.15 K) = -905.489 kJ/mol – S^o(298.15 K) = 50.053 J/mol-K – p^o = 0.1 MPa



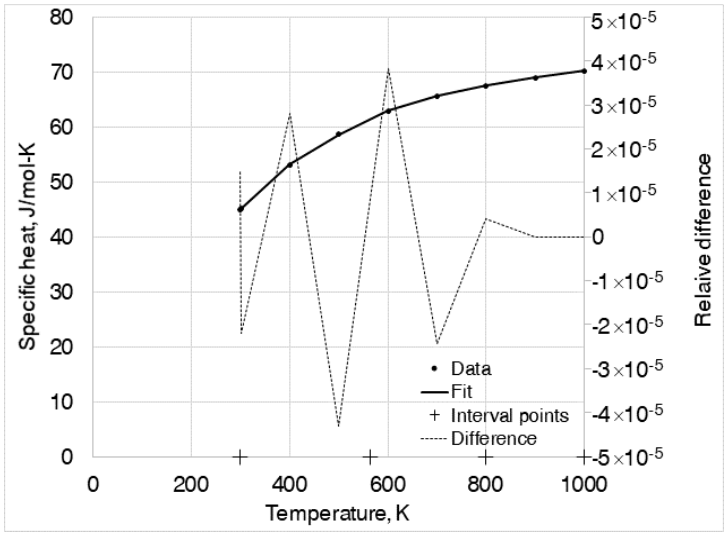


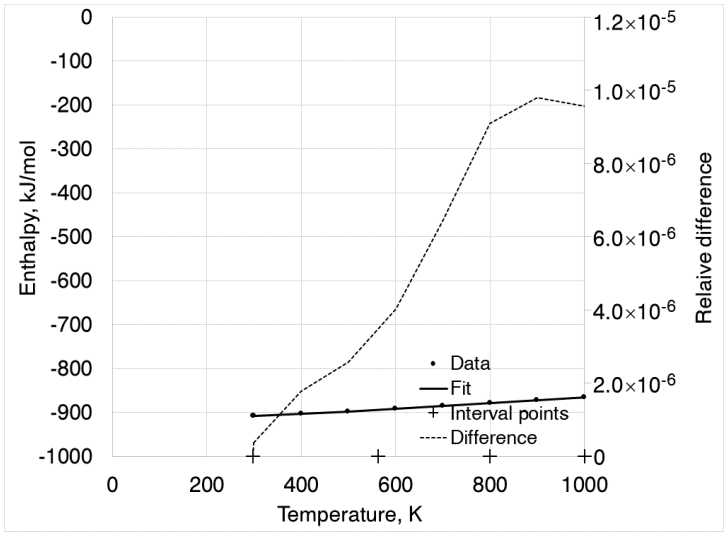


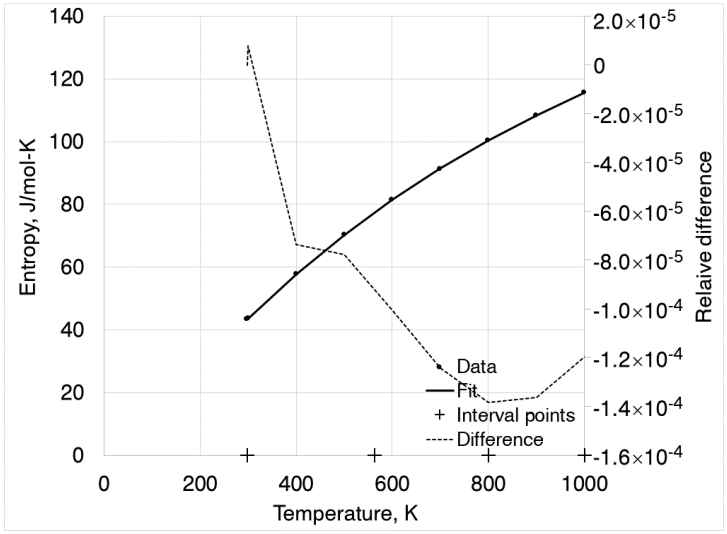


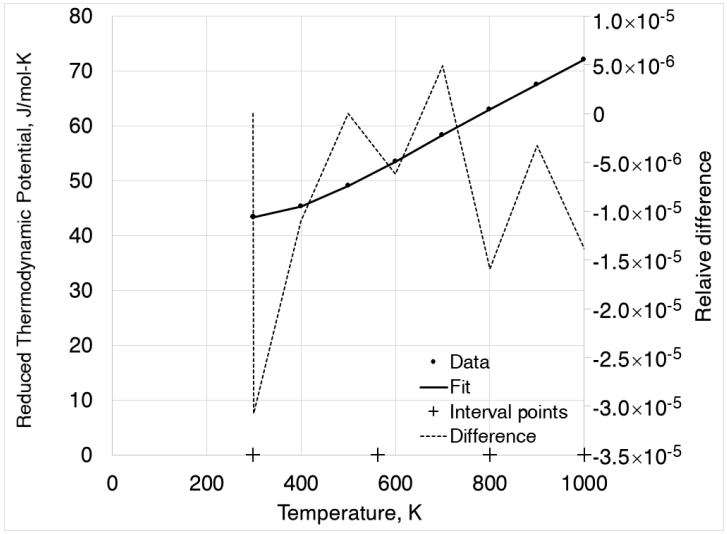
Silicon Oxide, Cristobalite, Low (SiO_2) – O2Si1(cr)_02

MW = 60.0843 g/mol – $\Delta_f H^o$(298.15 K) = -908.346 kJ/mol – S^o(298.15 K) = 43.396 J/mol-K – p^o = 0.1 MPa



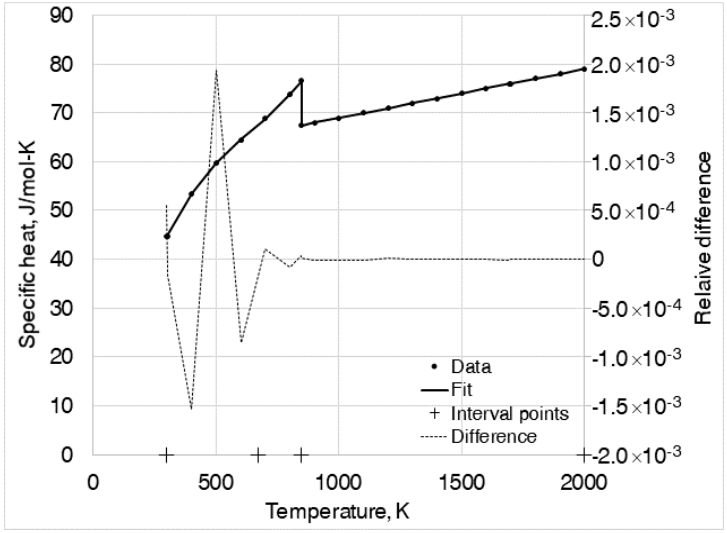


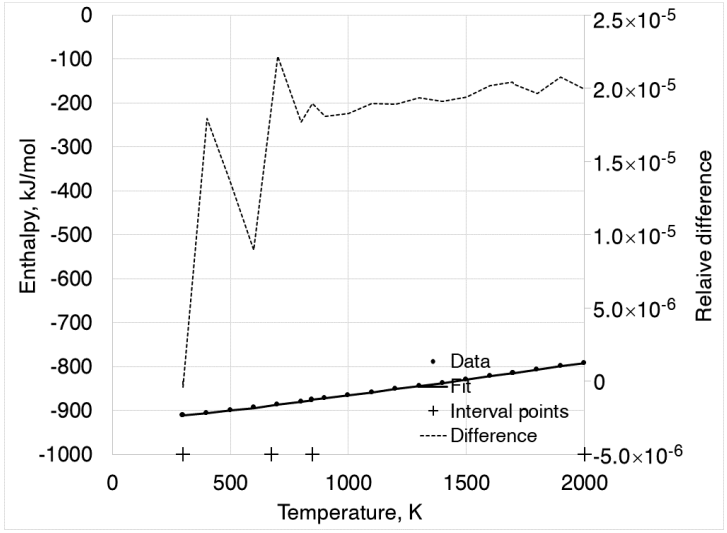


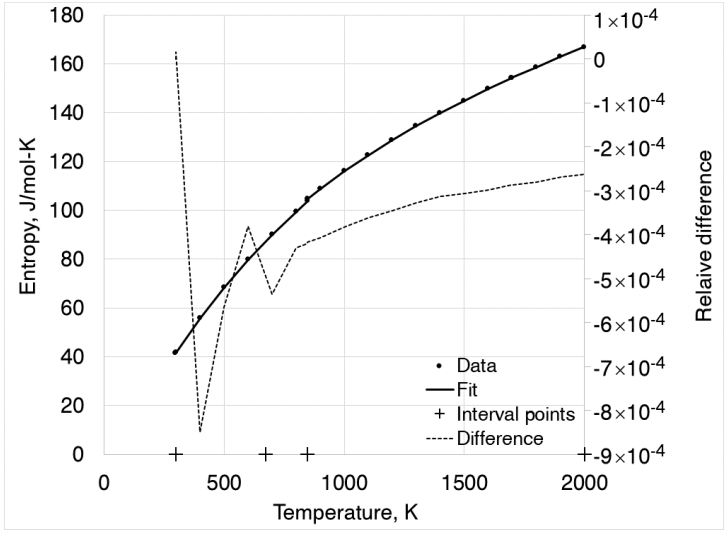


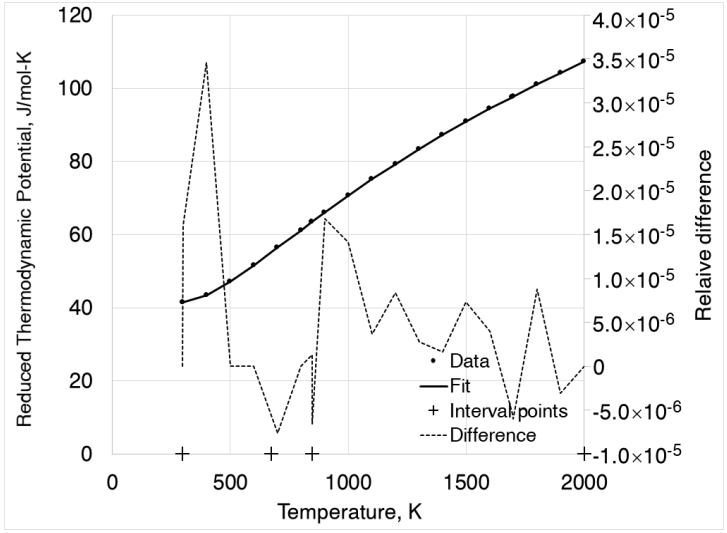
Silicon Oxide, Quartz (SiO_2) – O2Si1(cr)_03

MW = 60.0843 g/mol – $\Delta_f H^o$(298.15 K) = -910.857 kJ/mol – S^o(298.15 K) = 41.463 J/mol-K – p^o = 0.1 MPa



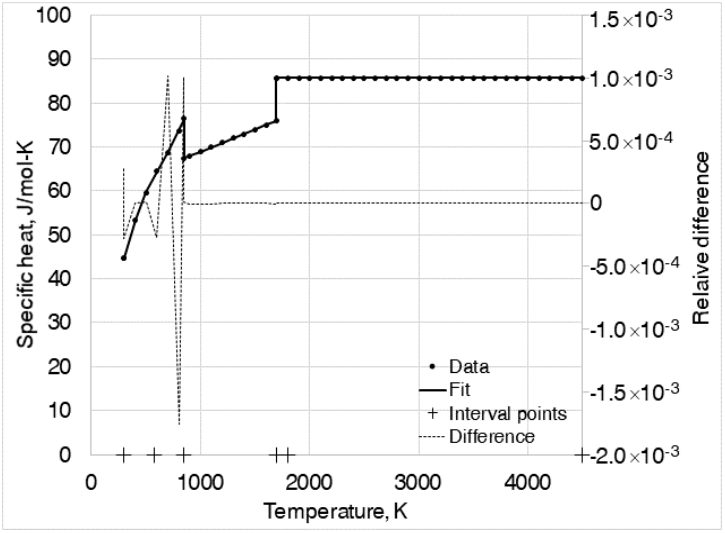


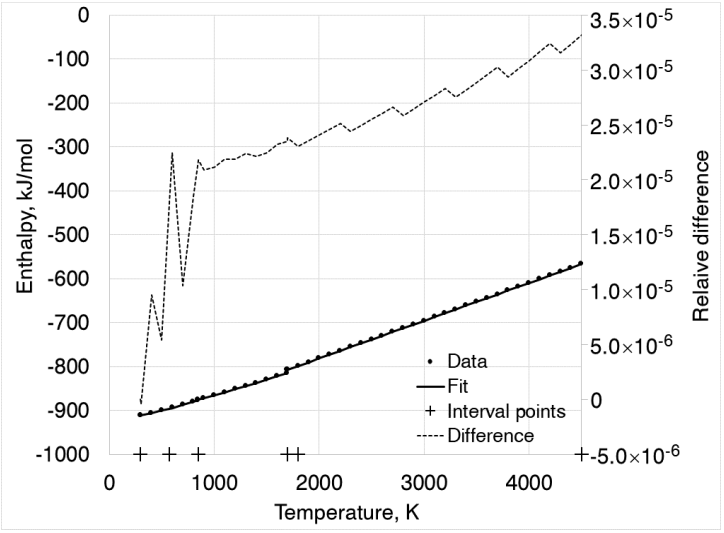


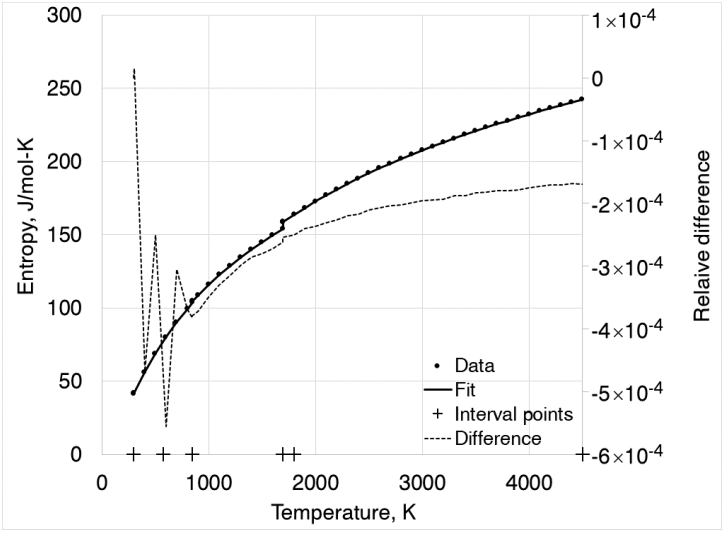


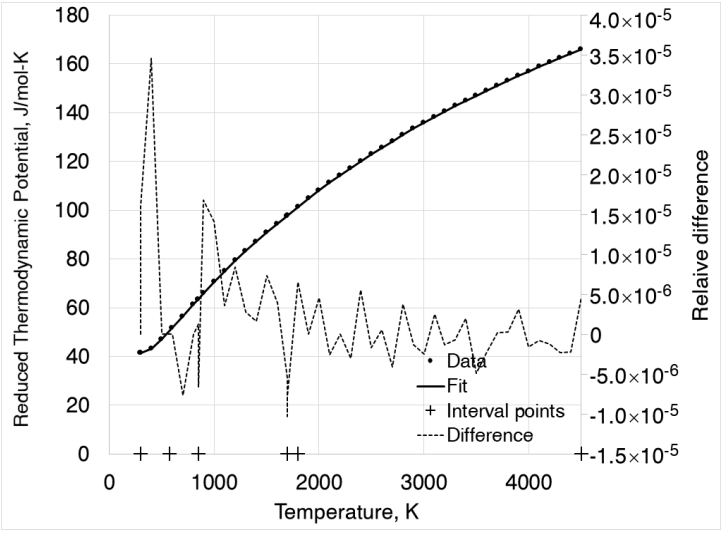
Silicon Oxide (SiO_2) – O2Si1(cr,l)

MW = 60.0843 g/mol – $\Delta_f H^o$(298.15 K) = -910.857 kJ/mol – S^o(298.15 K) = 41.463 J/mol-K – p^o = 0.1 MPa



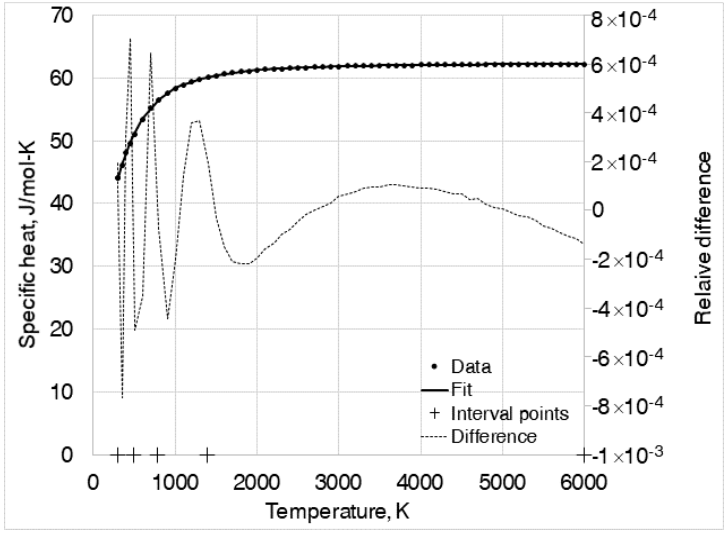


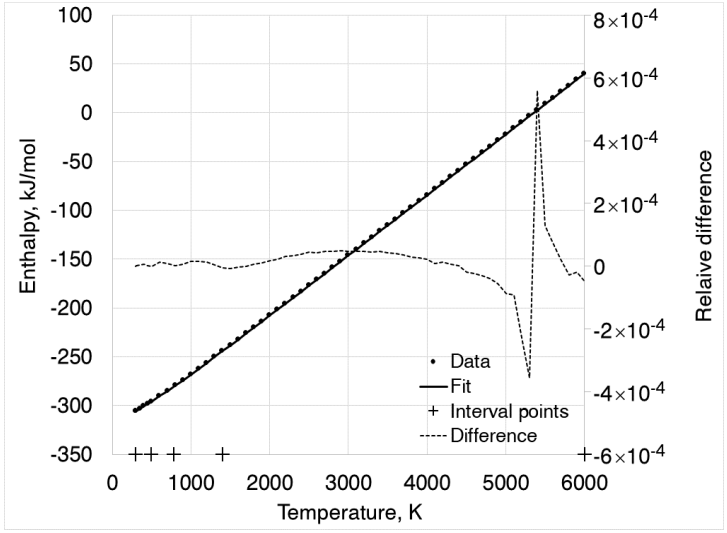


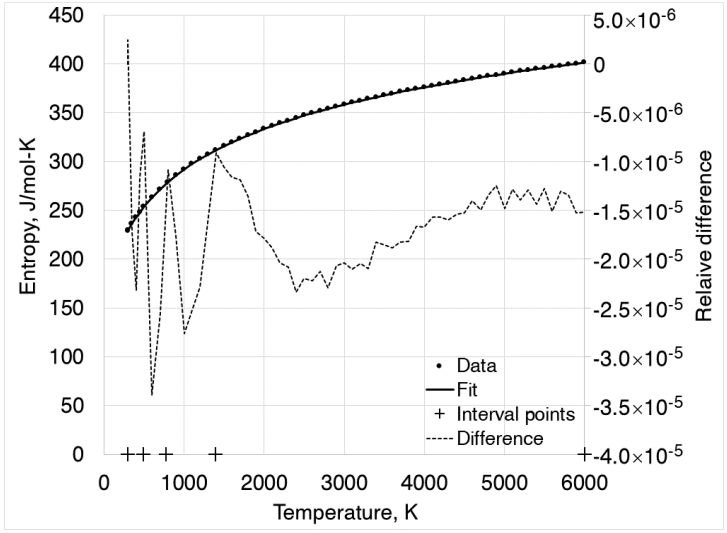


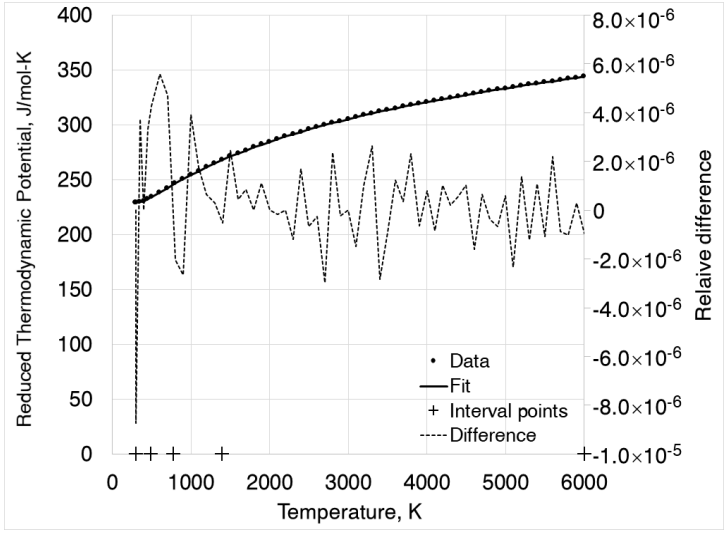
Silicon Oxide (SiO_2) – O2Si1(g)

MW = 60.0843 g/mol – $\Delta_f H^o$(298.15 K) = -305.432 kJ/mol – S^o(298.15 K) = 228.976 J/mol-K – p^o = 0.1 MPa



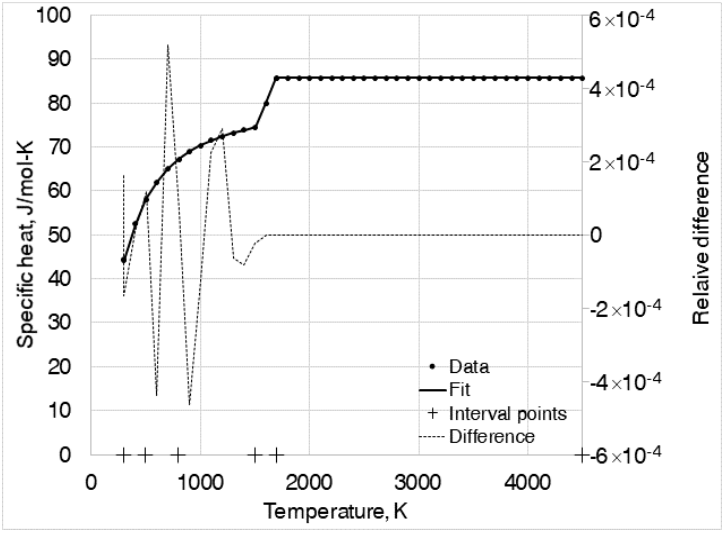


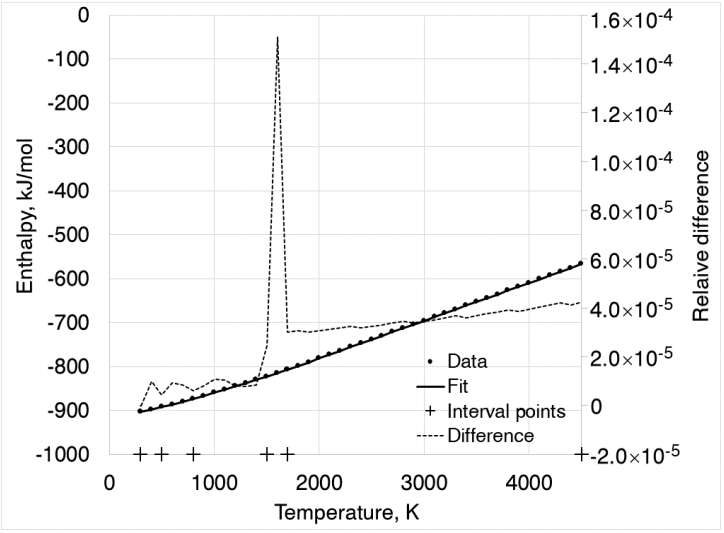


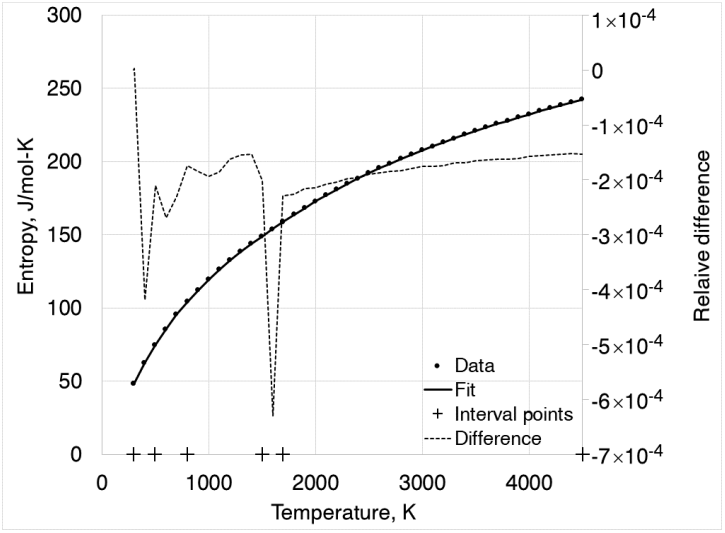


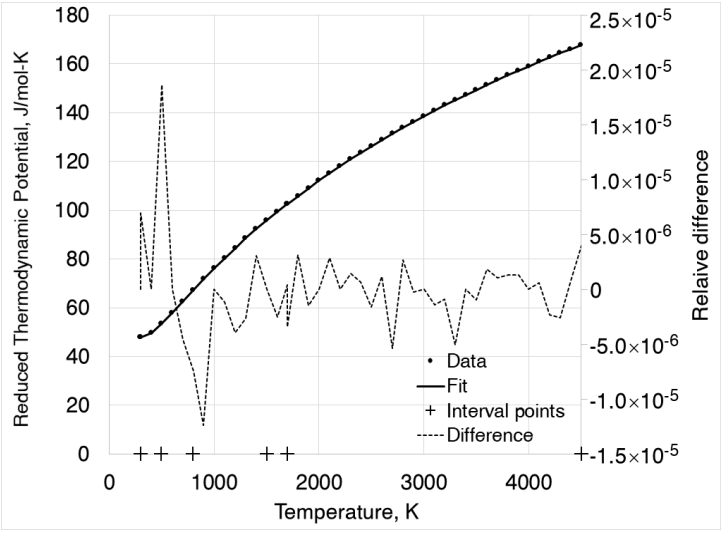
Silicon Oxide (SiO_2) – O2Si1(l)

MW = 60.0843 g/mol – $\Delta_f H^o$(298.15 K) = -902.661 kJ/mol – S^o(298.15 K) = 47.928 J/mol-K – p^o = 0.1 MPa



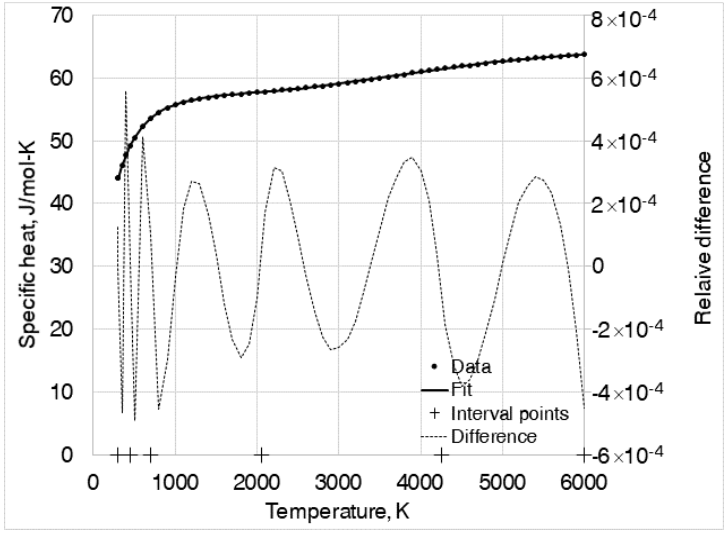


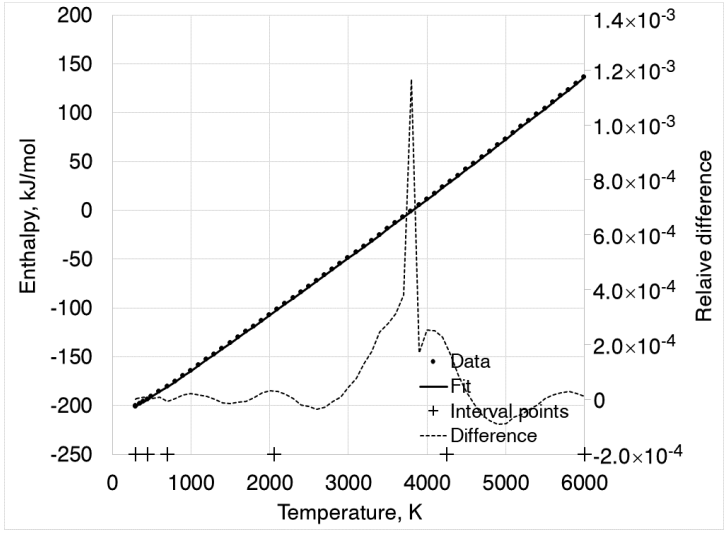


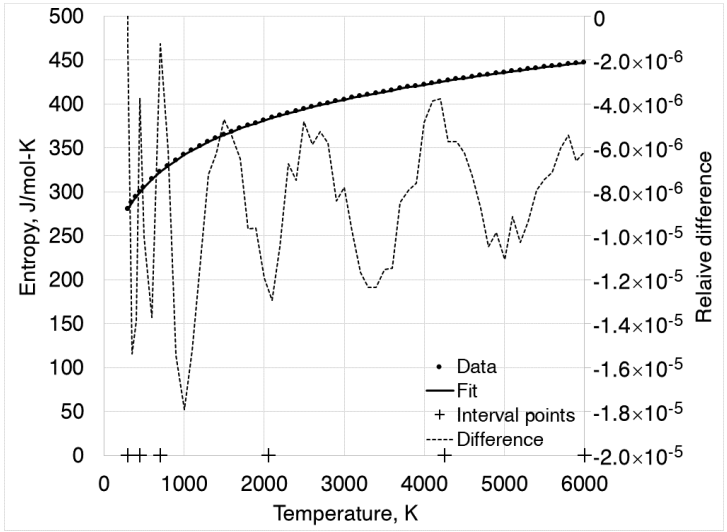


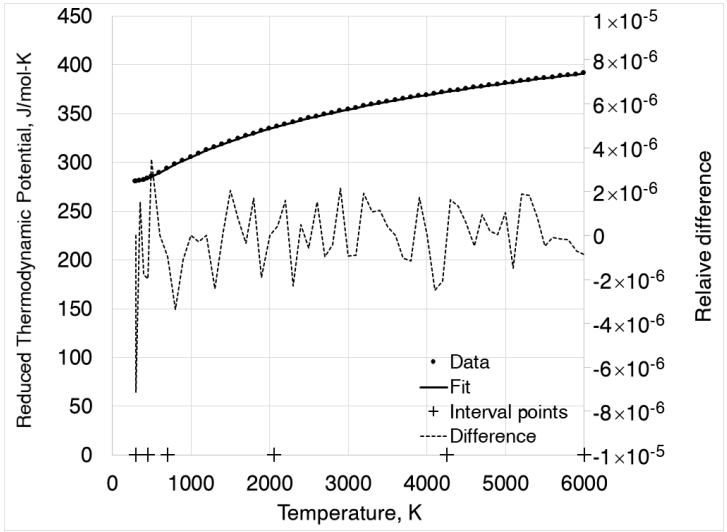
Tantalum Oxide (TaO_2) – O2Ta1(g)

MW = 212.9467 g/mol – $\Delta_f H^o$(298.15 K) = -200.832 kJ/mol – S^o(298.15 K) = 280.306 J/mol-K – p^o = 0.1 MPa



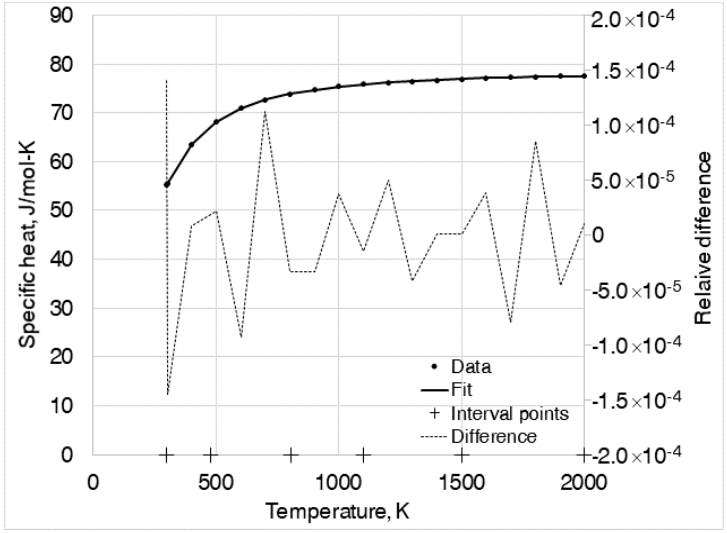


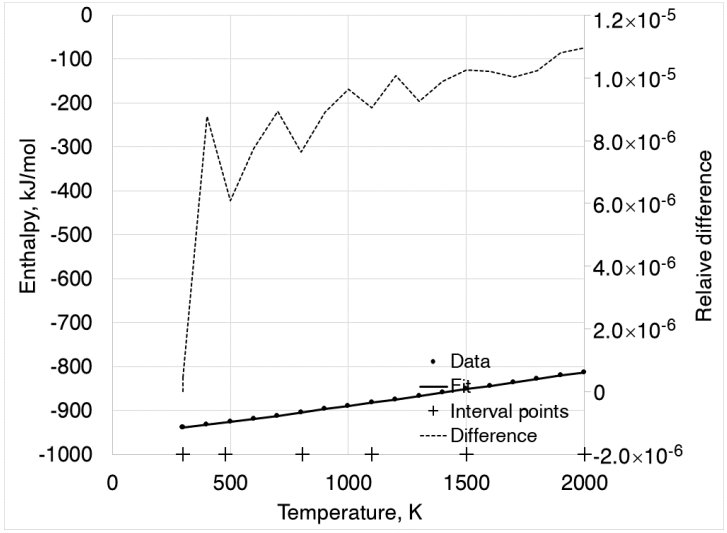


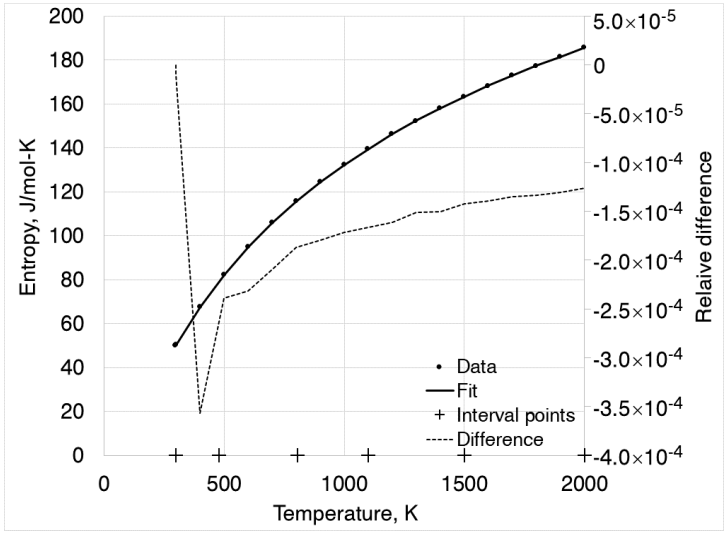


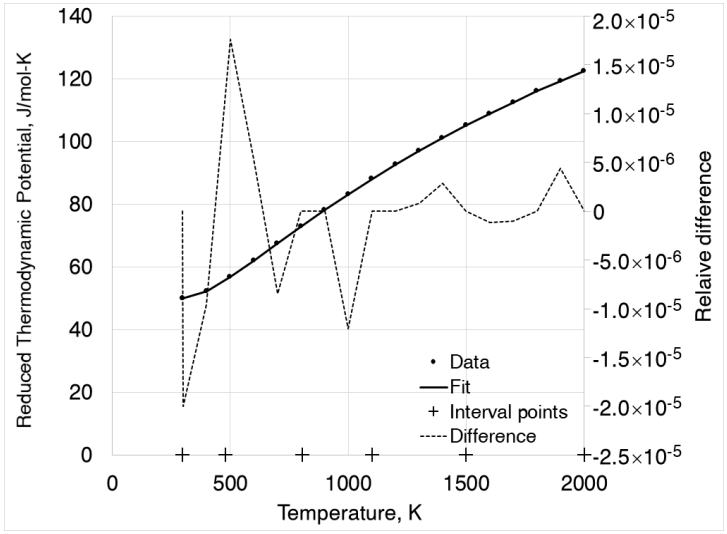
Titanium Oxide, Anatase (TiO_2) – O2Ti1(cr)_01

MW = 79.88 g/mol – $\Delta_f H^o$(298.15 K) = -938.722 kJ/mol – S^o(298.15 K) = 49.907 J/mol-K – p^o = 0.1 MPa



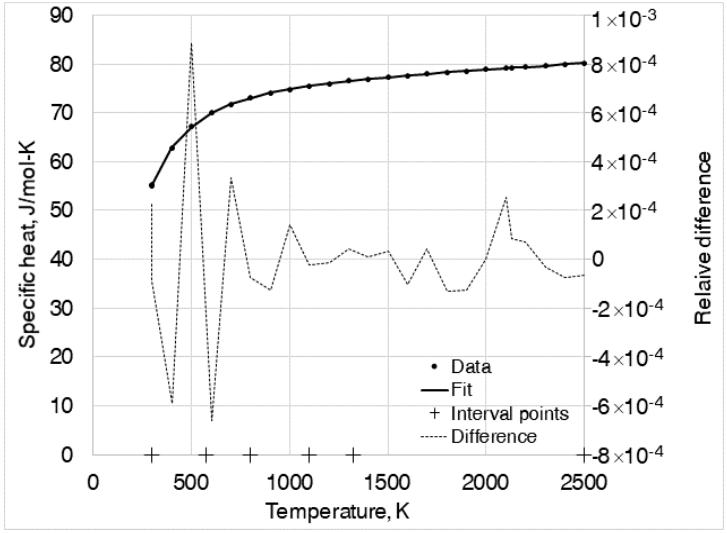


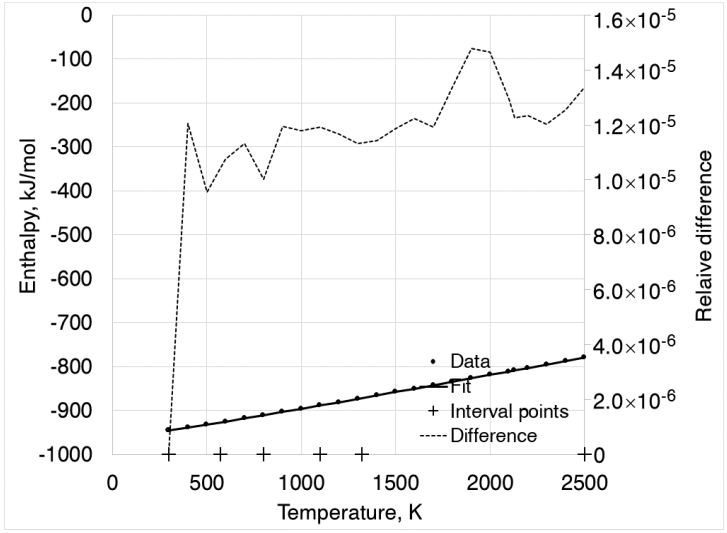


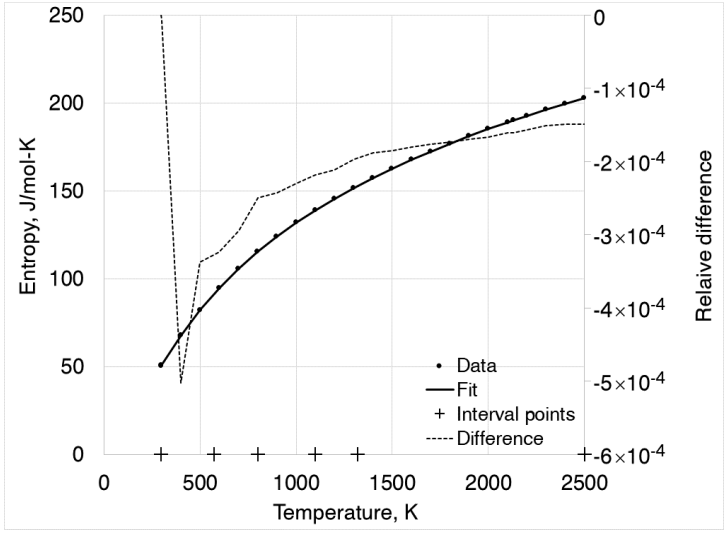


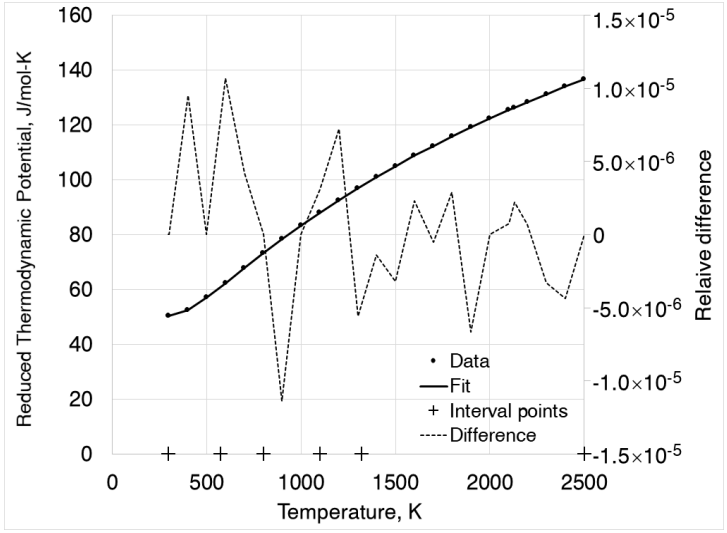
Titanium Oxide, Rutile (TiO_2) – O2Ti1(cr)_02

MW = 79.88 g/mol – $\Delta_f H^o$(298.15 K) = -944.747 kJ/mol – S^o(298.15 K) = 50.292 J/mol-K – p^o = 0.1 MPa



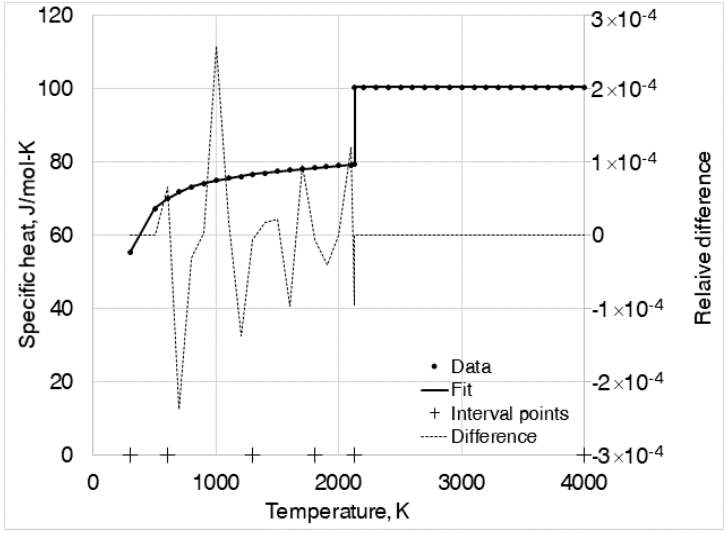


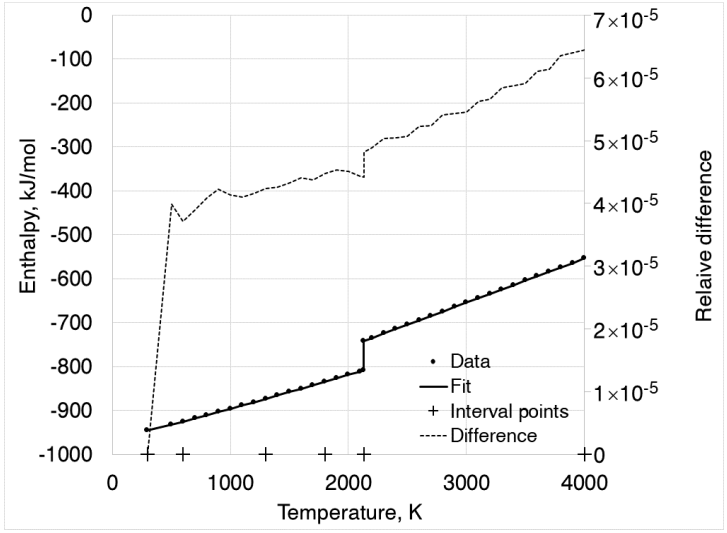


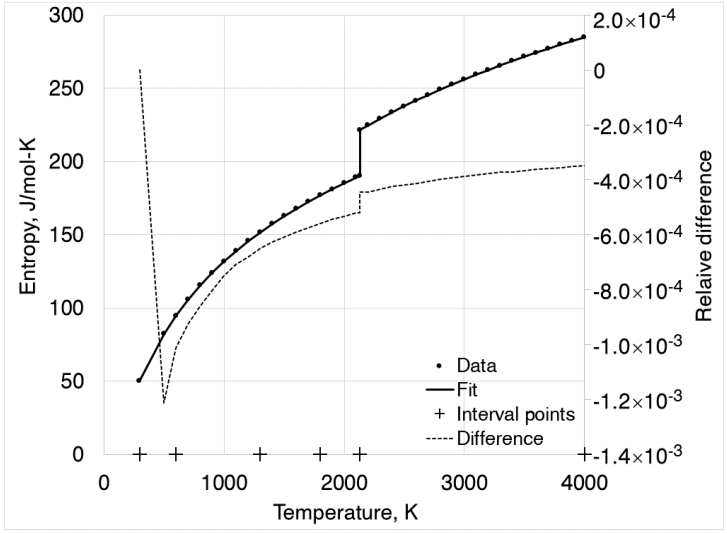


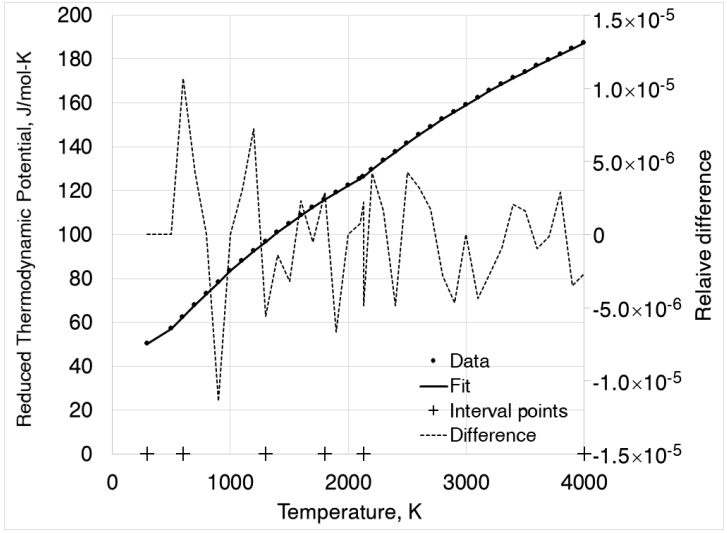
Titanium Oxide (TiO_2) – O2Ti1(cr,l)

MW = 79.88 g/mol – $\Delta_f H^o$(298.15 K) = -944.747 kJ/mol – S^o(298.15 K) = 50.292 J/mol-K – p^o = 0.1 MPa



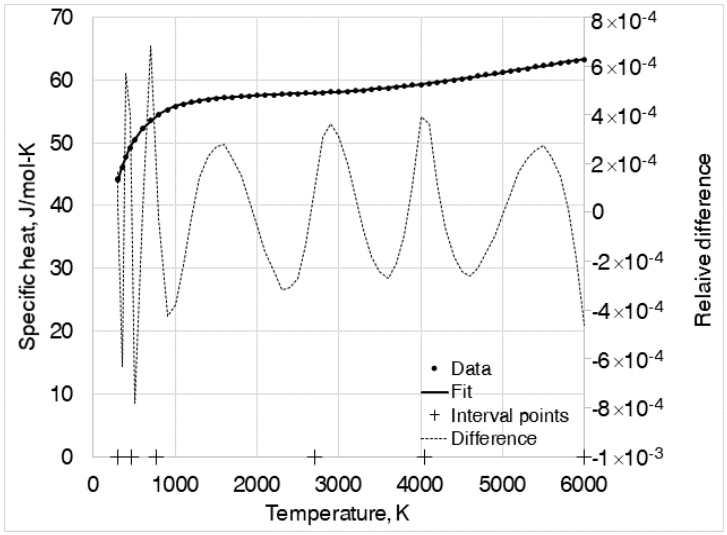


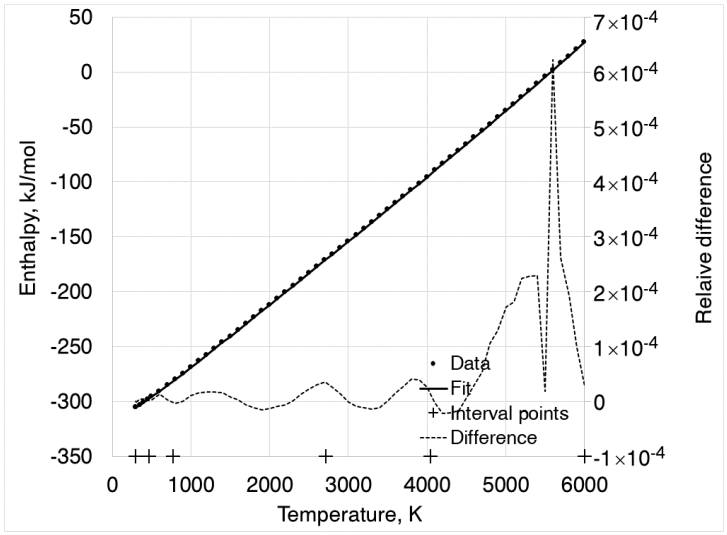


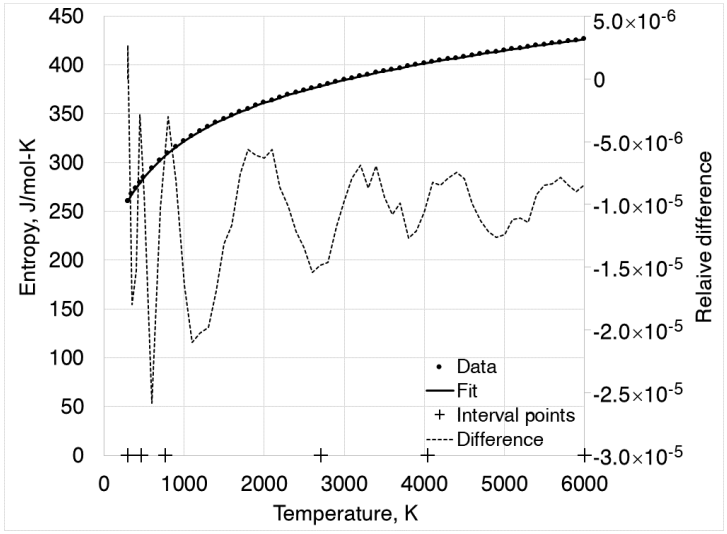


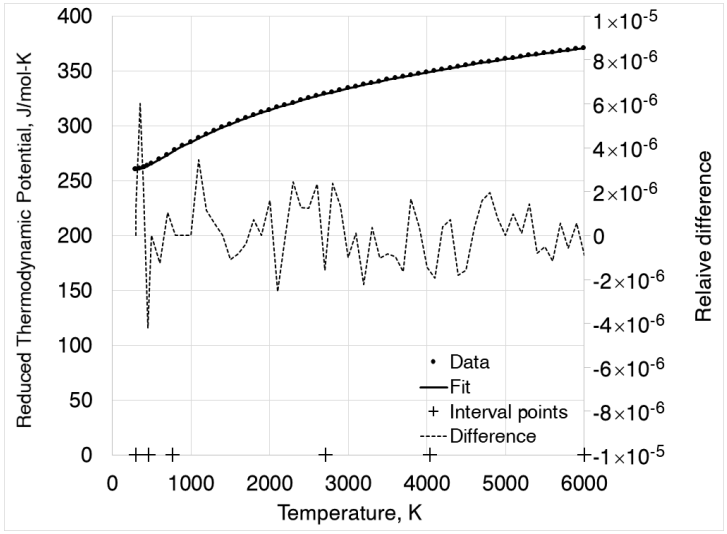
Titanium Oxide (TiO_2) – O2Ti1(g)

MW = 79.88 g/mol – $\Delta_f H^o$(298.15 K) = -305.432 kJ/mol – S^o(298.15 K) = 260.148 J/mol-K – p^o = 0.1 MPa



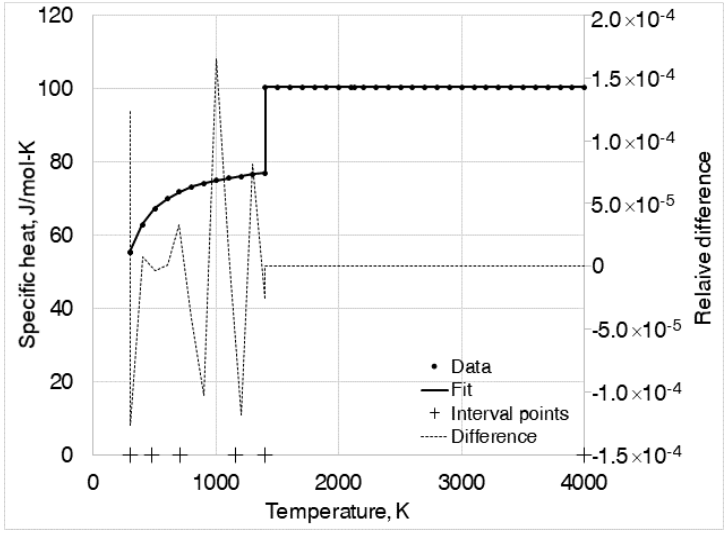


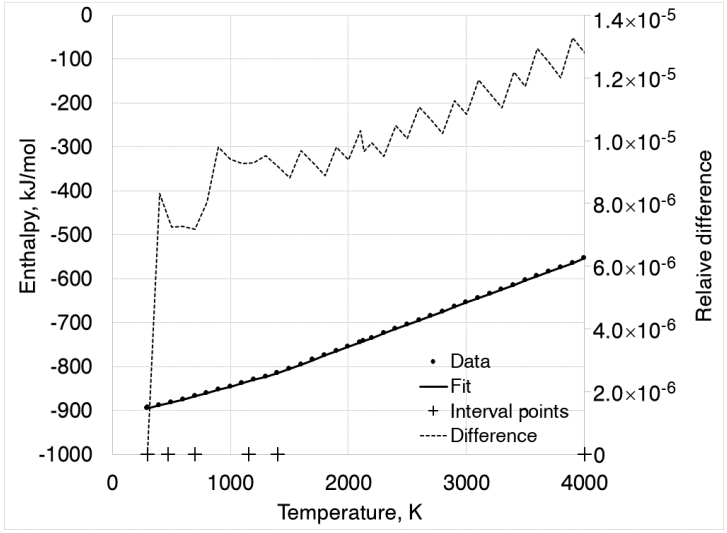


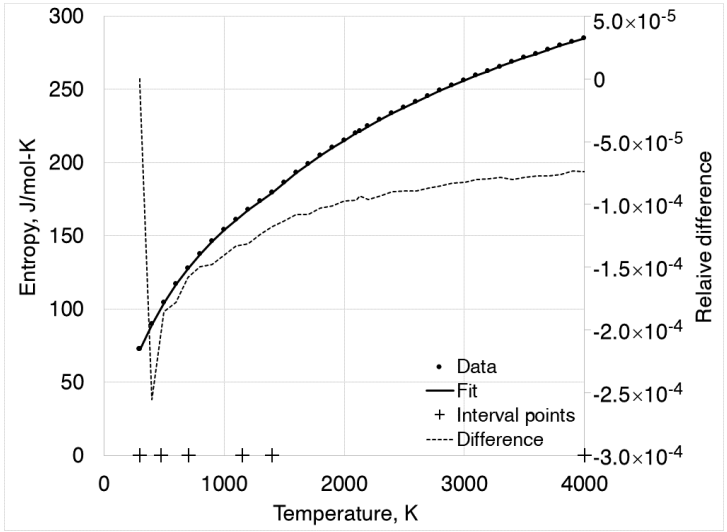


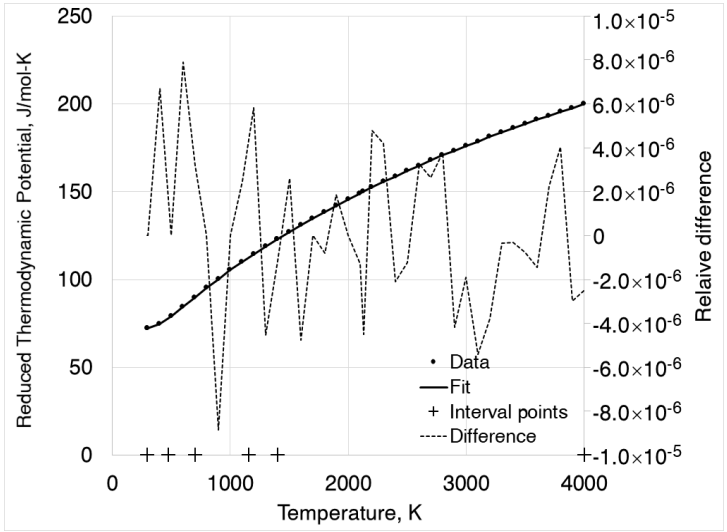
Titanium Oxide (TiO_2) – O2Ti1(l)

MW = 79.88 g/mol – $\Delta_f H^o$(298.15 K) = -894.055 kJ/mol – S^o(298.15 K) = 72.345 J/mol-K – p^o = 0.1 MPa



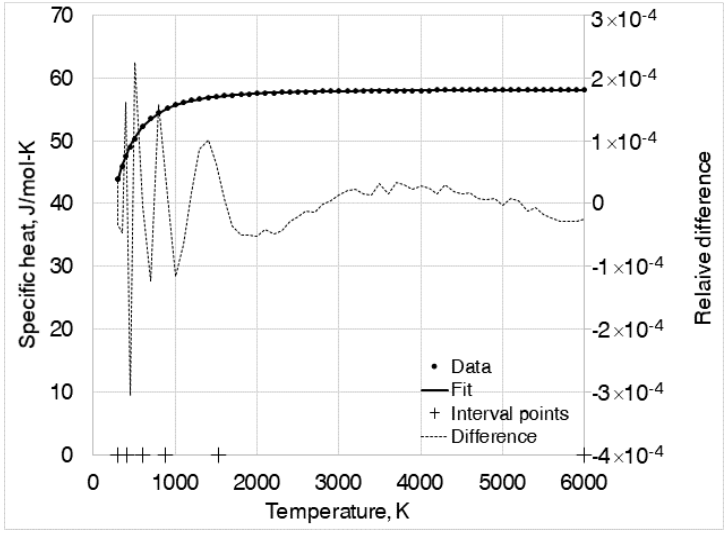


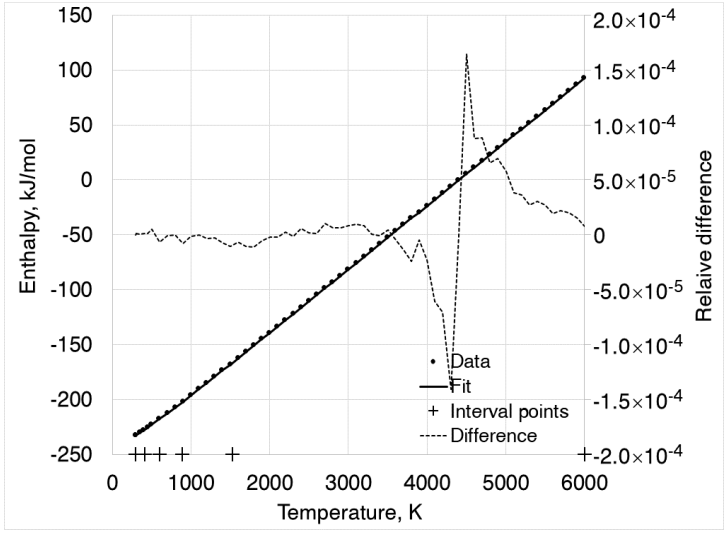


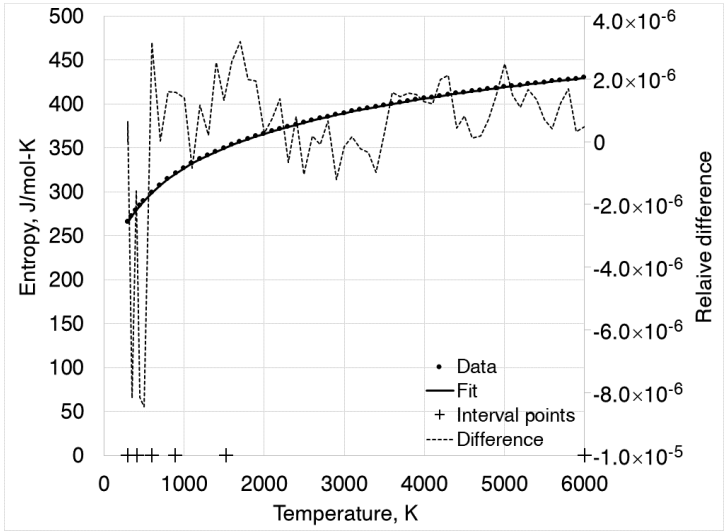


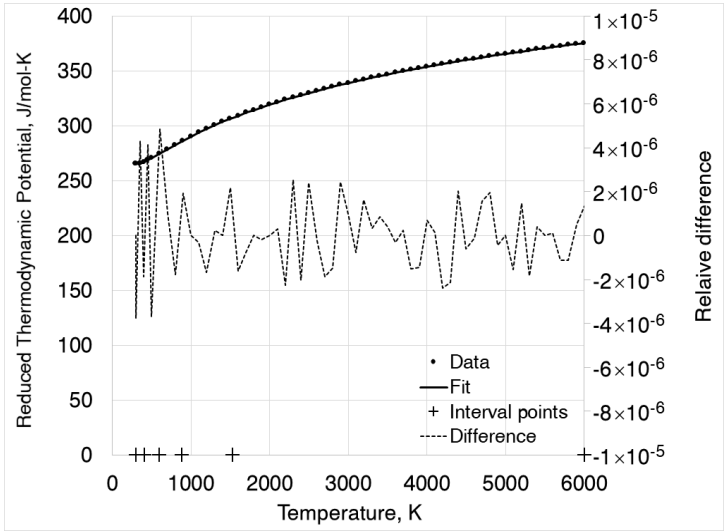
Vanadium Oxide (VO_2) – O2V1(g)

MW = 82.9403 g/mol – $\Delta_f H^o$(298.15 K) = -232.63 kJ/mol – S^o(298.15 K) = 265.269 J/mol-K – p^o = 0.1 MPa



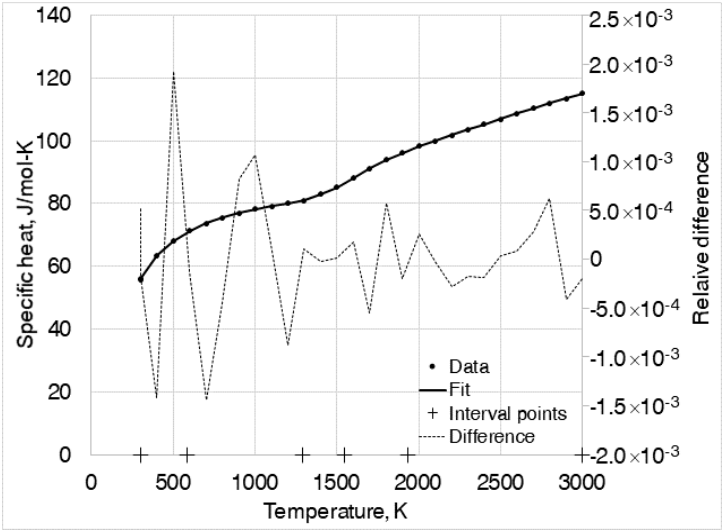


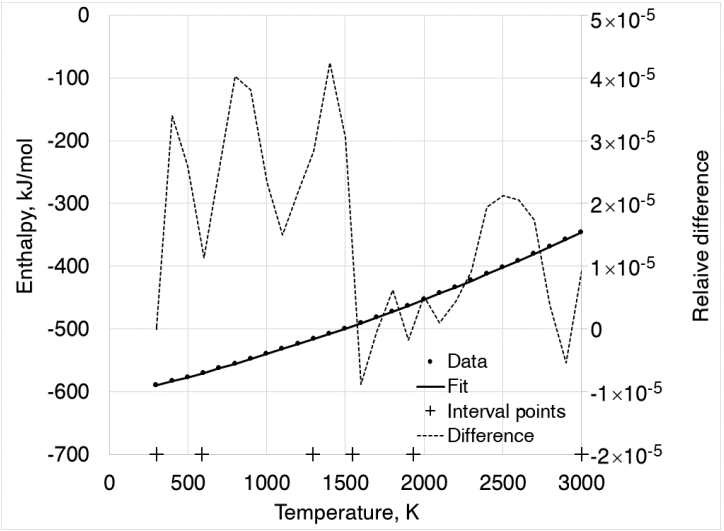


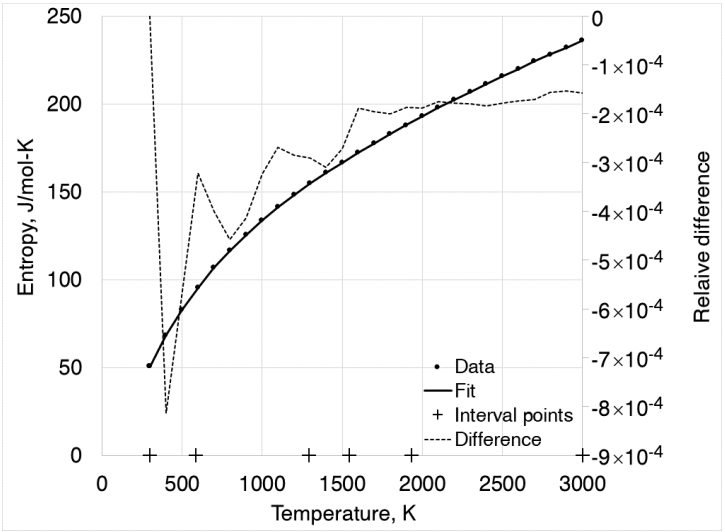


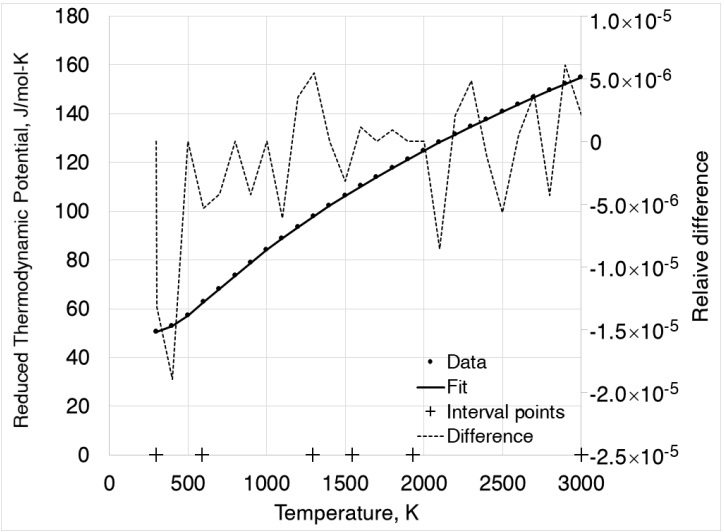
Tungsten Oxide (WO_2) – O2W1(cr)

MW = 215.85 g/mol – $\Delta_f H^o$(298.15 K) = -589.693 kJ/mol – S^o(298.15 K) = 50.528 J/mol-K – p^o = 0.1 MPa



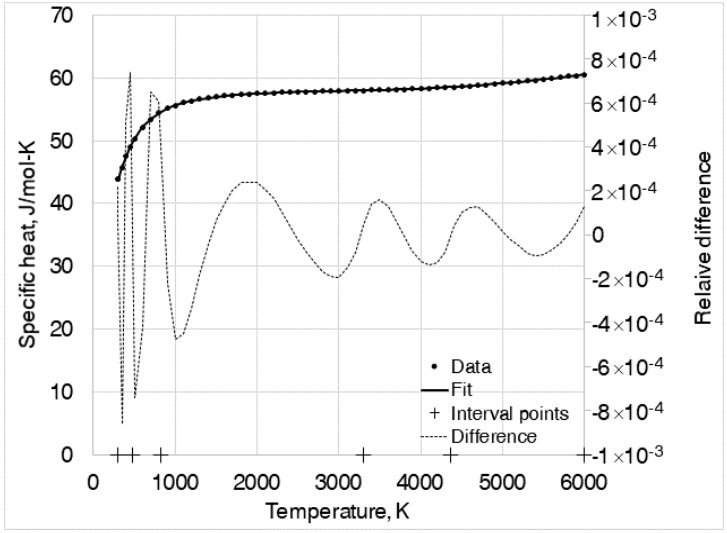


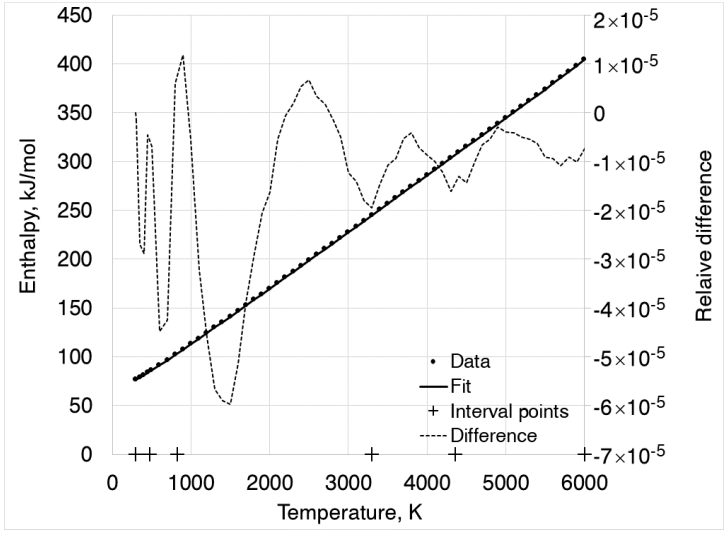


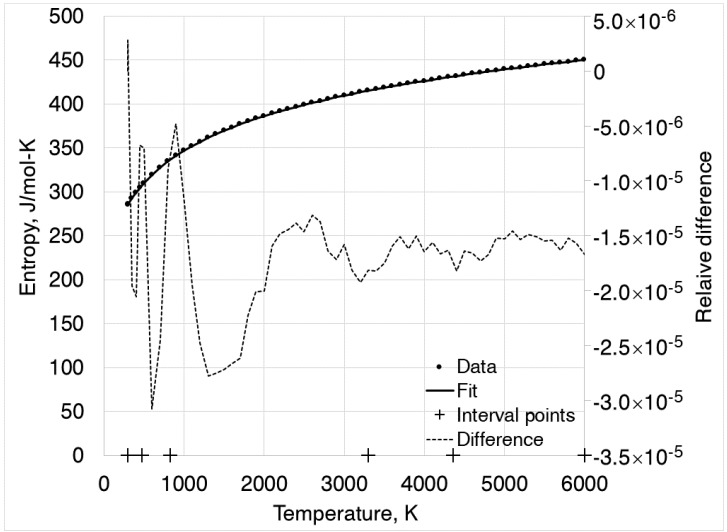


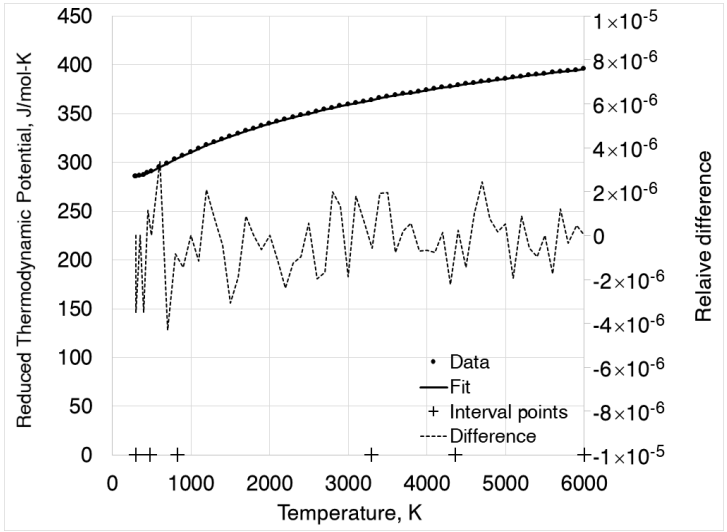
Tungsten Oxide (WO_2) – O2W1(g)

MW = 215.85 g/mol – $\Delta_f H^o$(298.15 K) = 76.567 kJ/mol – S^o(298.15 K) = 285.499 J/mol-K – p^o = 0.1 MPa



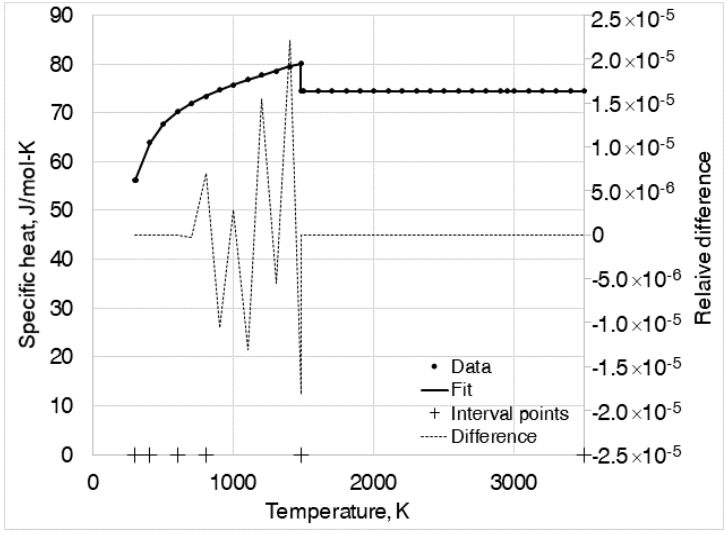


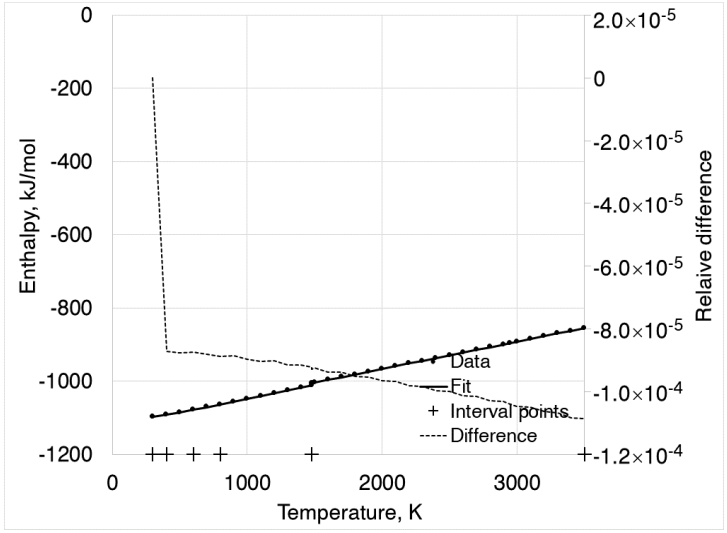


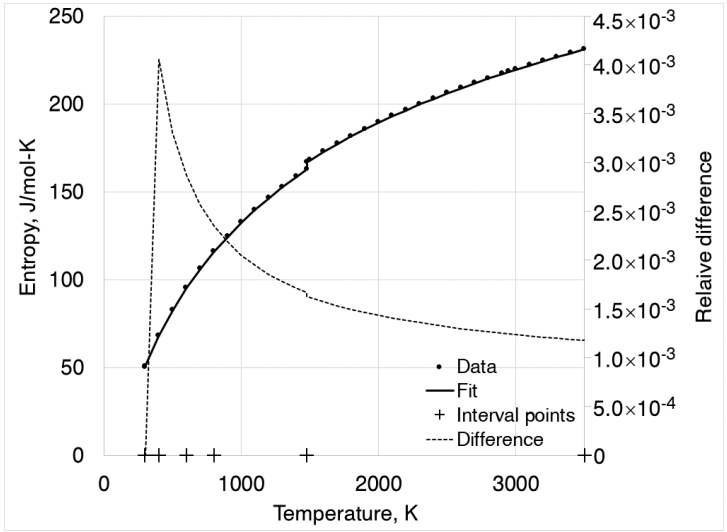


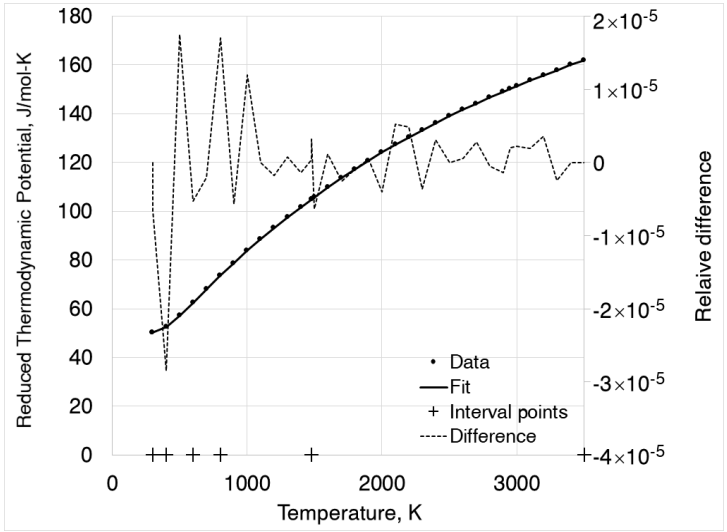
Zirconium Oxide (ZrO_2) – O2Zr1(cr)

MW = 123.22 g/mol – $\Delta_f H^o$(298.15 K) = -1097.463 kJ/mol – S^o(298.15 K) = 50.359 J/mol-K – p^o = 0.1 MPa



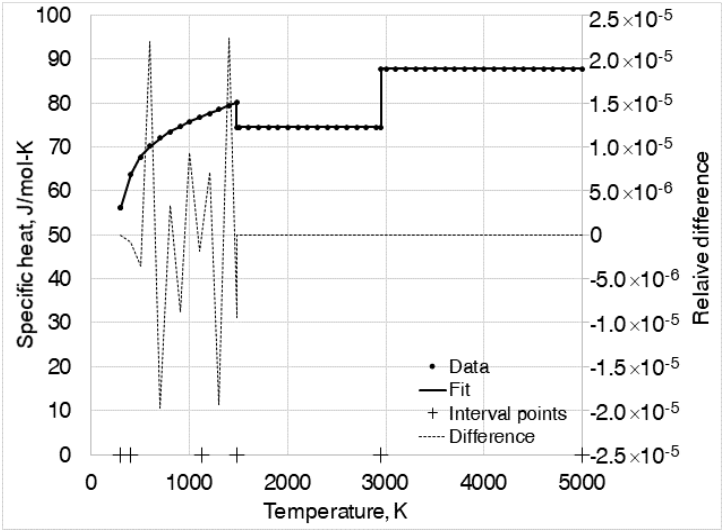


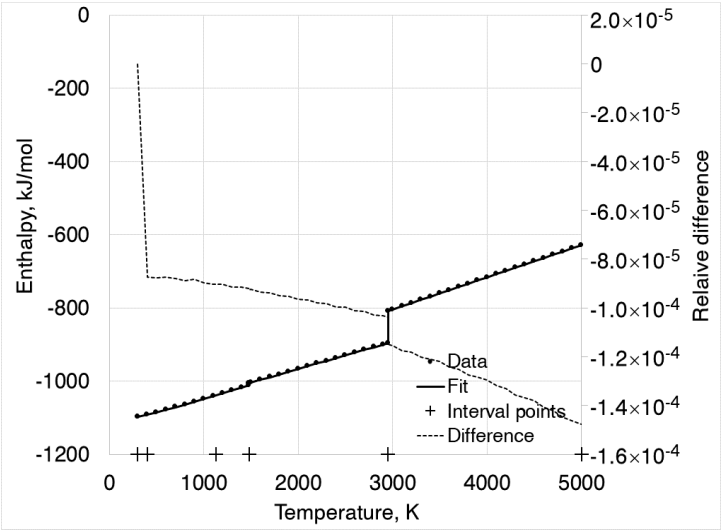


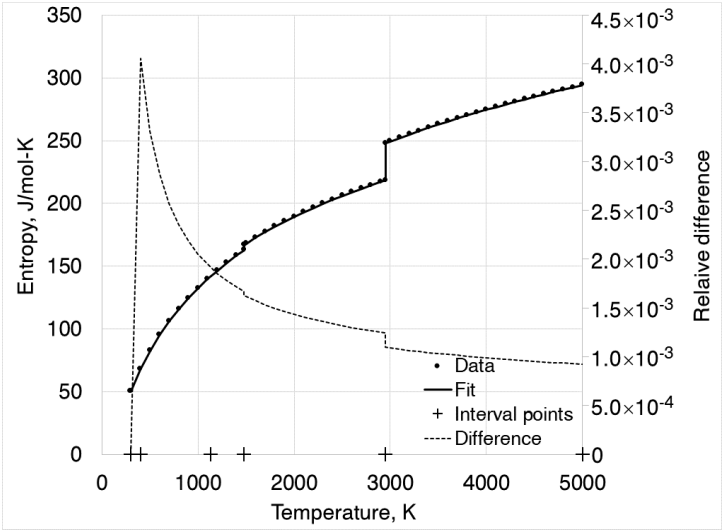


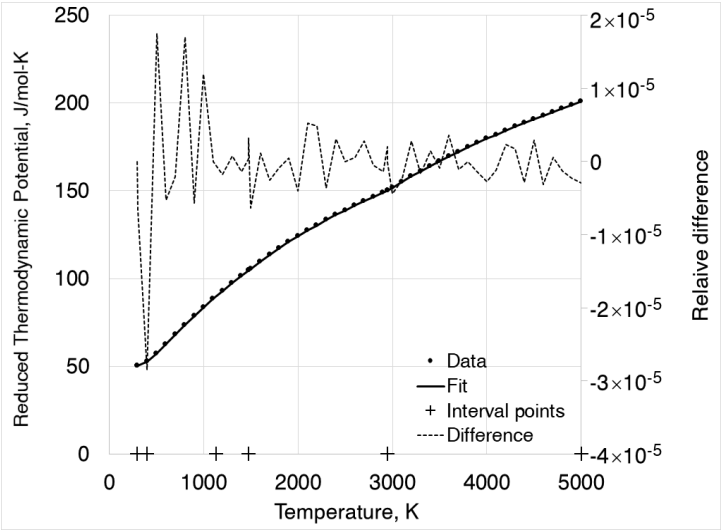
Zirconium Oxide (ZrO_2) – O2Zr1(cr,l)

MW = 123.22 g/mol – $\Delta_f H^o$(298.15 K) = -1097.463 kJ/mol – S^o(298.15 K) = 50.359 J/mol-K – p^o = 0.1 MPa



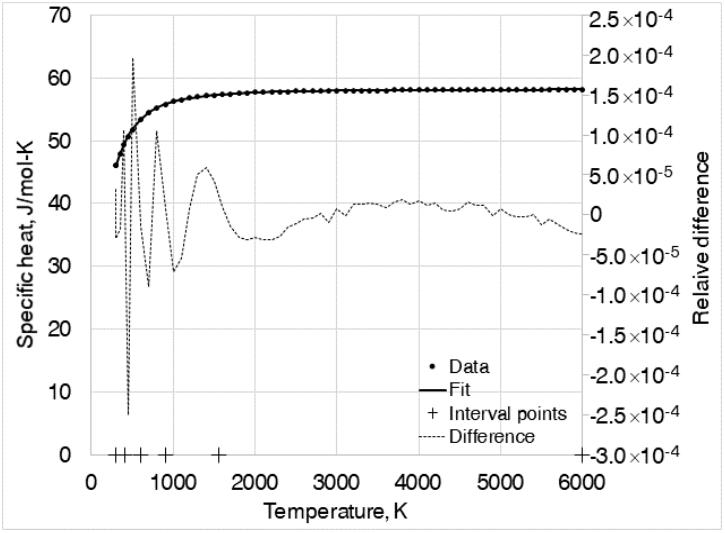


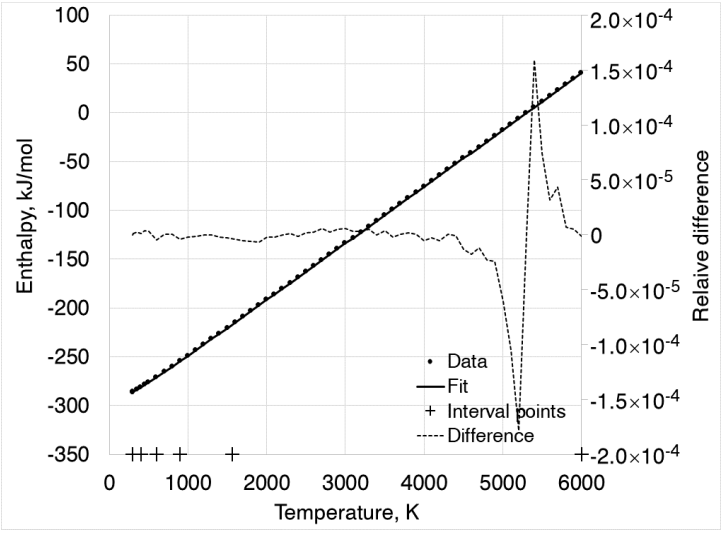


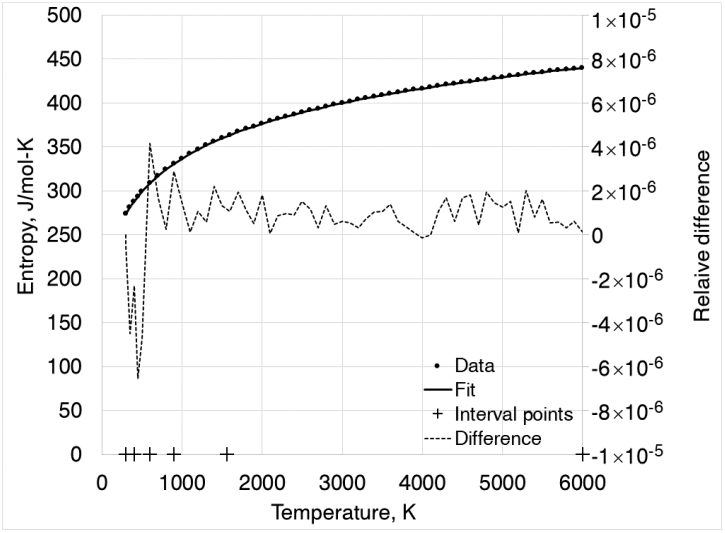


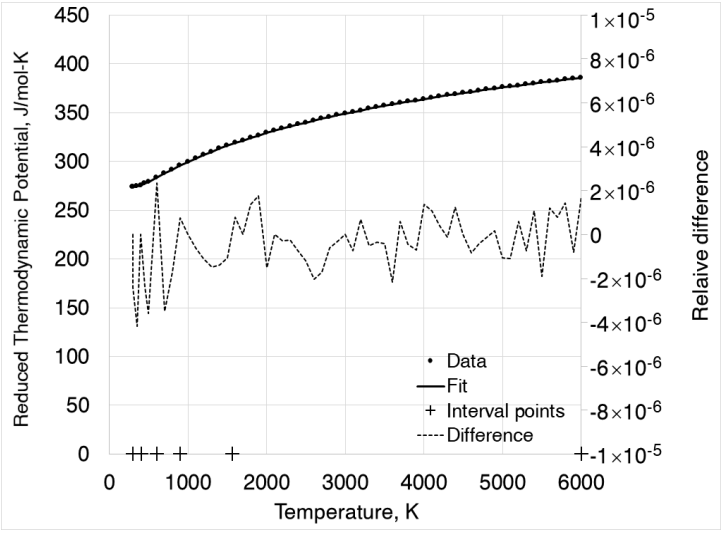
Zirconium Oxide (ZrO_2) – O2Zr1(g)

MW = 123.22 g/mol – $\Delta_f H^o$(298.15 K) = -286.186 kJ/mol – S^o(298.15 K) = 273.746 J/mol-K – p^o = 0.1 MPa



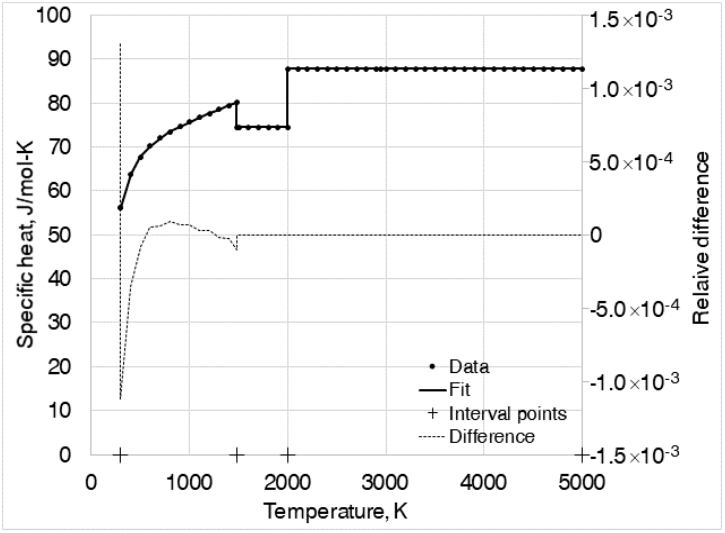


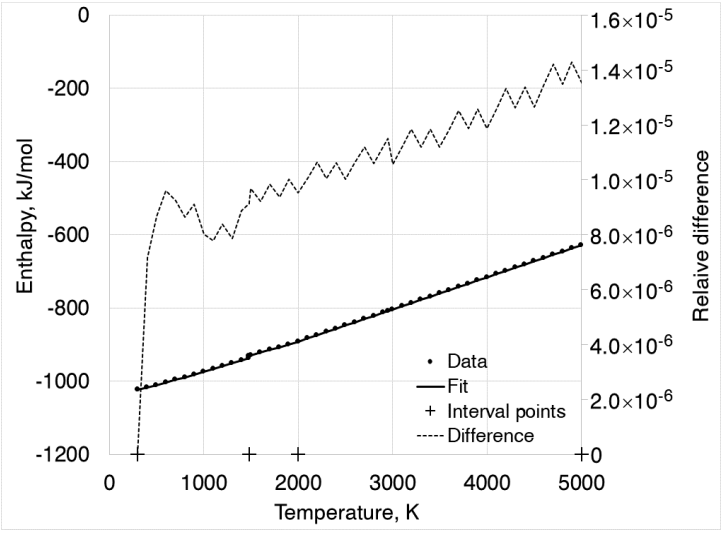


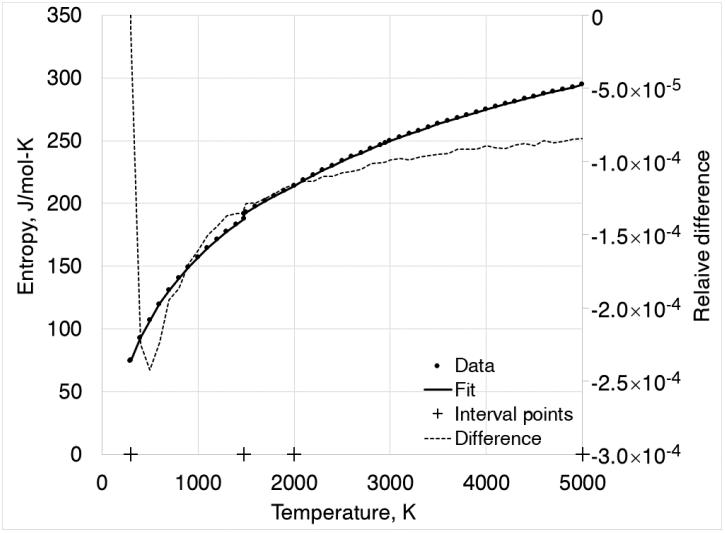


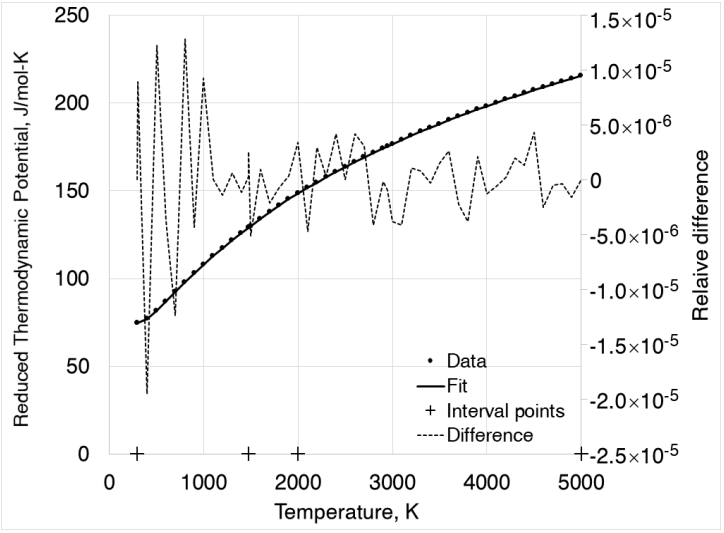
Zirconium Oxide (ZrO_2) – O2Zr1(l)

MW = 123.22 g/mol – $\Delta_f H^o$(298.15 K) = -1023.155 kJ/mol – S^o(298.15 K) = 74.657 J/mol-K – p^o = 0.1 MPa



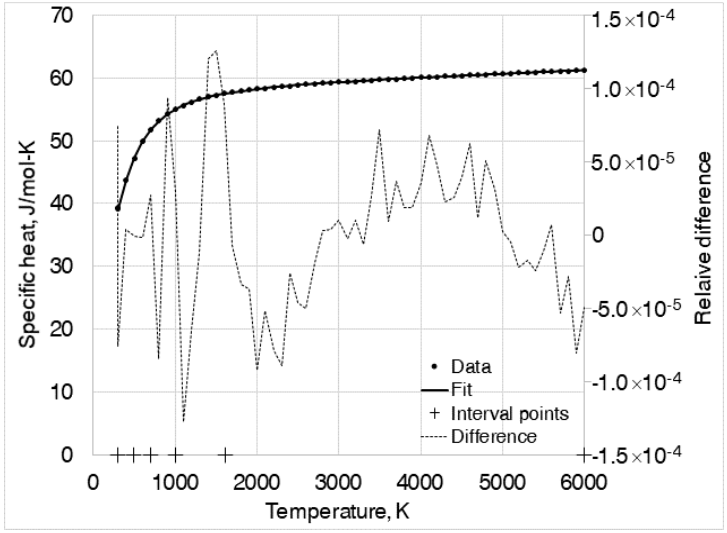


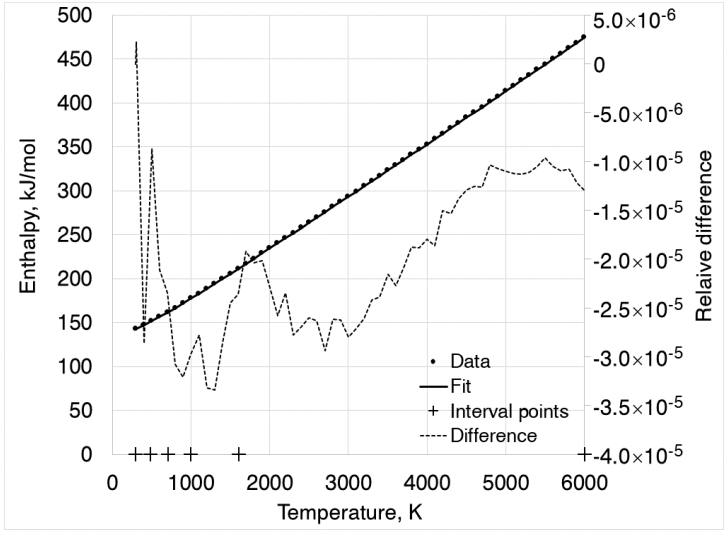


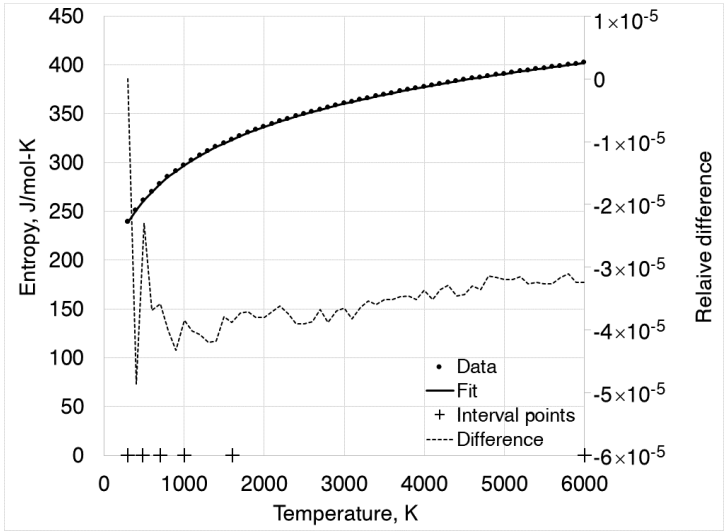


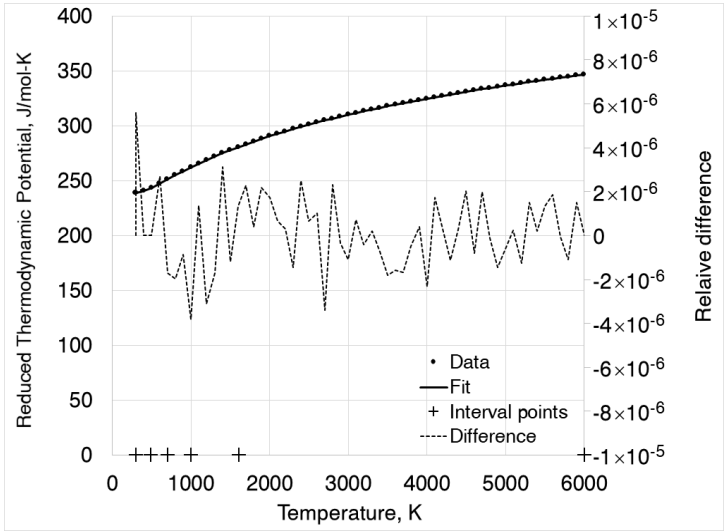
Ozone (O_3) – O3(g)

MW = 47.9982 g/mol – $\Delta_f H^o$(298.15 K) = 142.674 kJ/mol – S^o(298.15 K) = 238.932 J/mol-K – p^o = 0.1 MPa



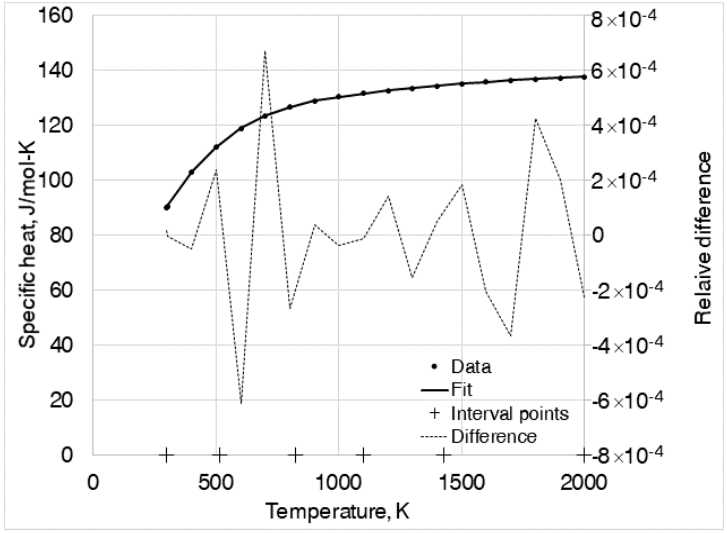


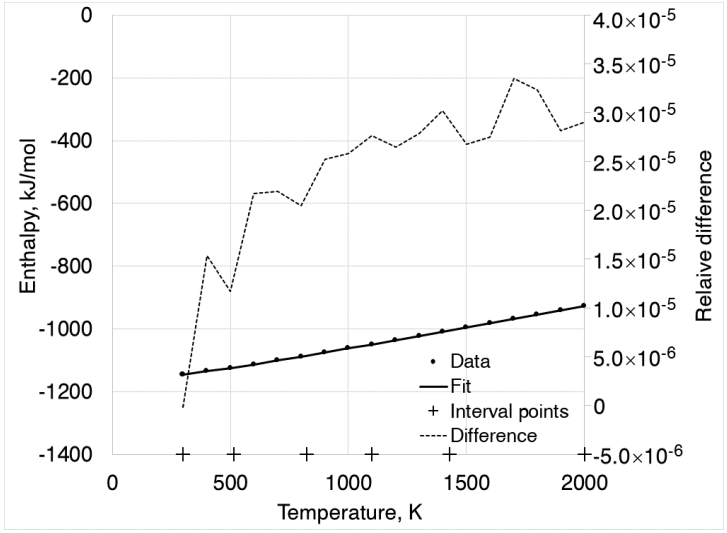


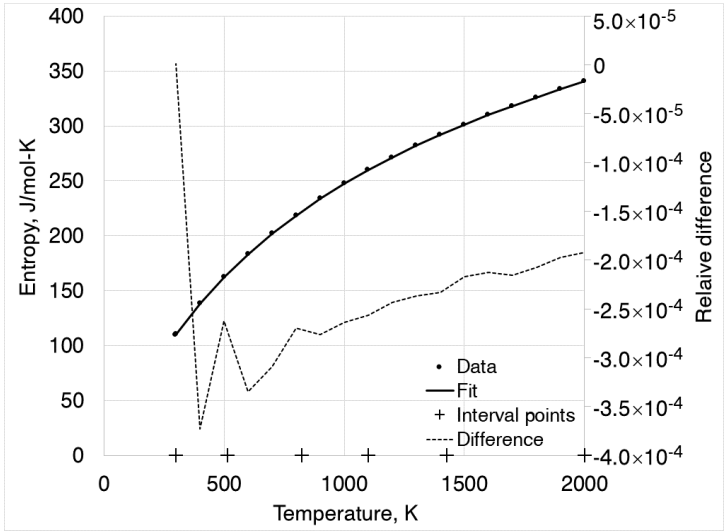


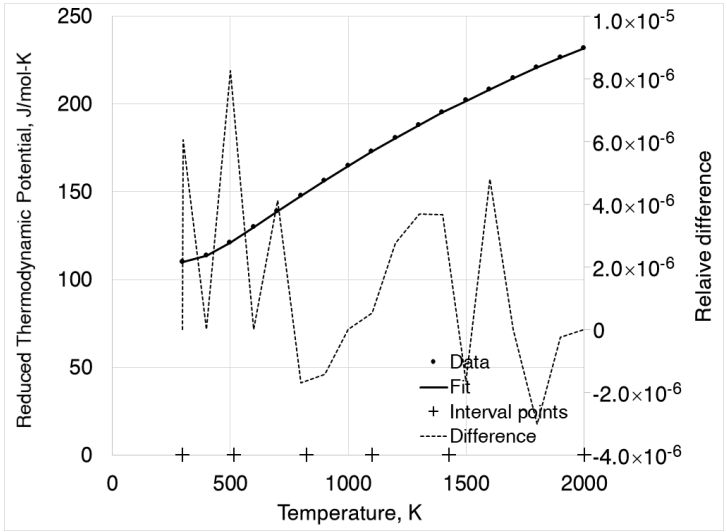
Lead Silicate ($PbSiO_3$) – O3Pb1Si1(cr)

MW = 283.3 g/mol – $\Delta_f H^o$(298.15 K) = -1144.993 kJ/mol – S^o(298.15 K) = 109.927 J/mol-K – p^o = 0.1 MPa



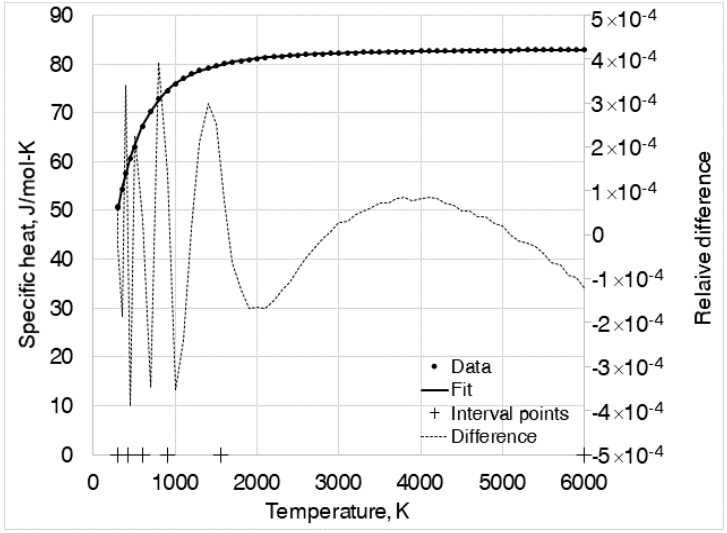


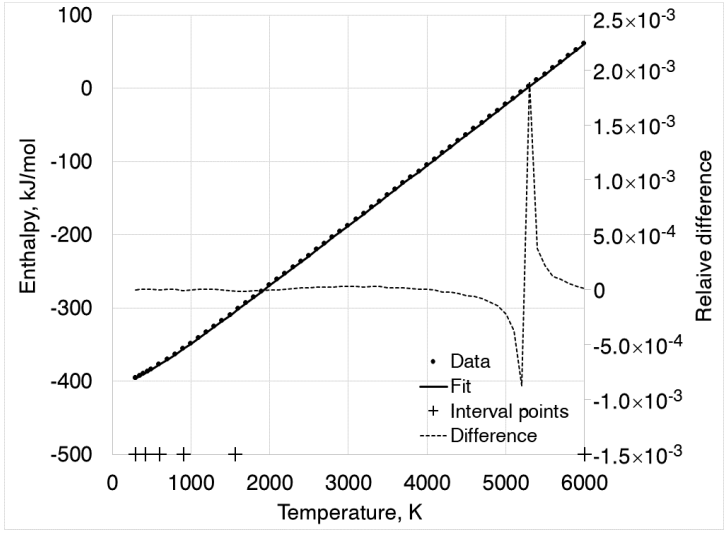


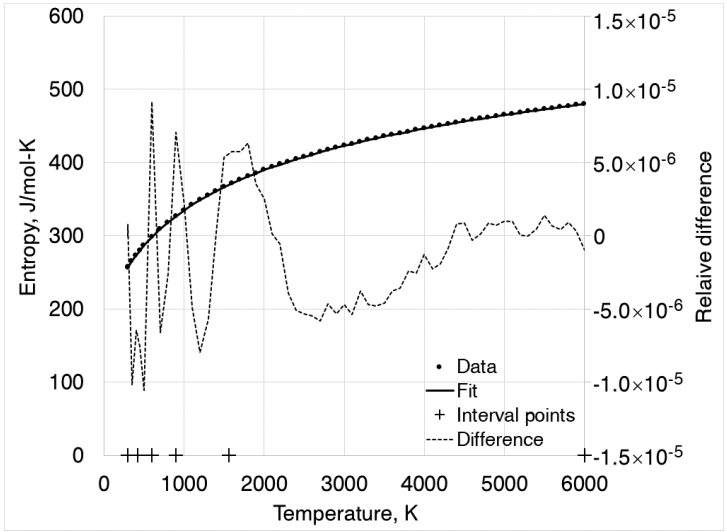


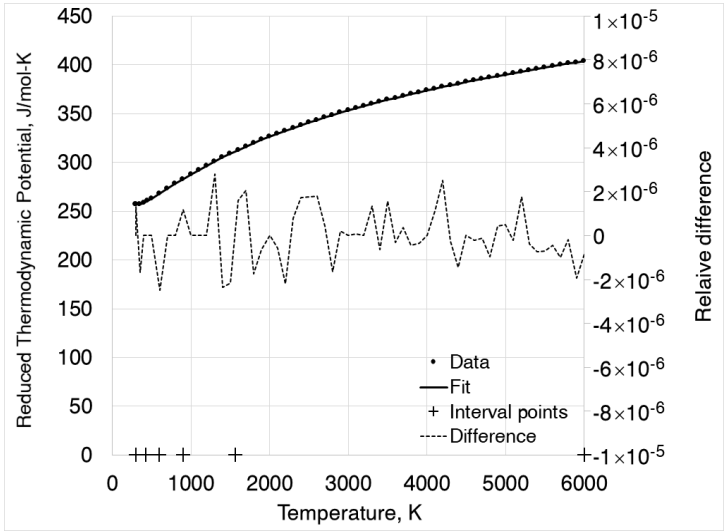
Sulfur Trioxide (SO_3) – O3S1(g)

MW = 80.06 g/mol – $\Delta_f H^o$(298.15 K) = -395.765 kJ/mol – S^o(298.15 K) = 256.769 J/mol-K – p^o = 0.1 MPa



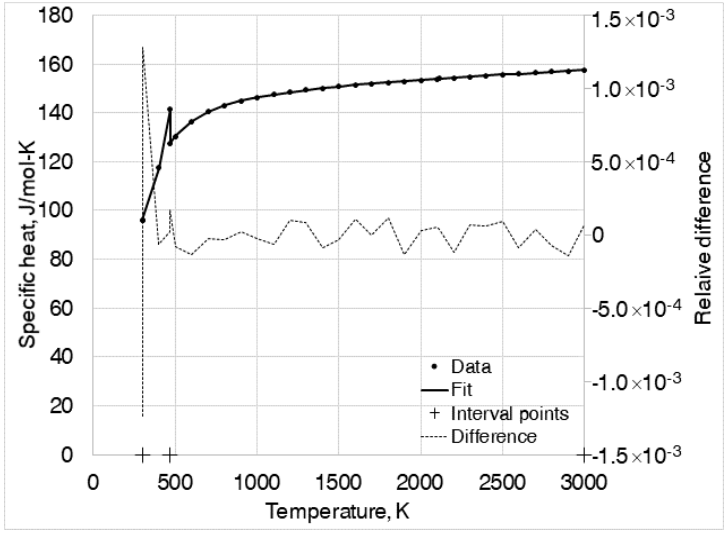


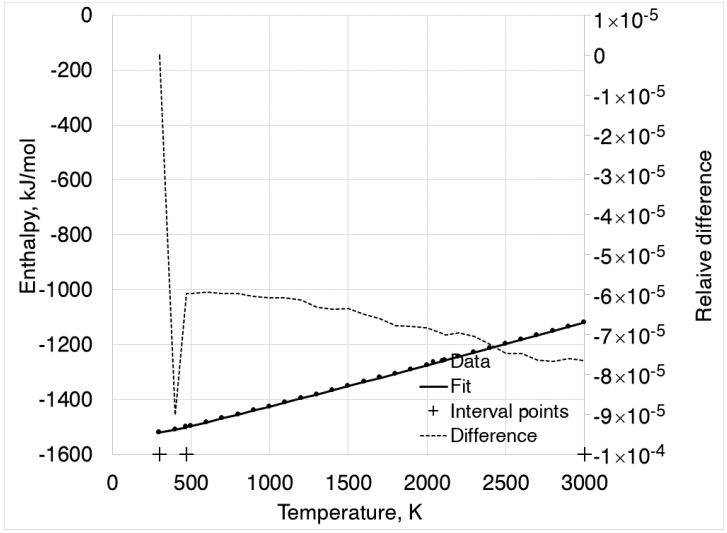


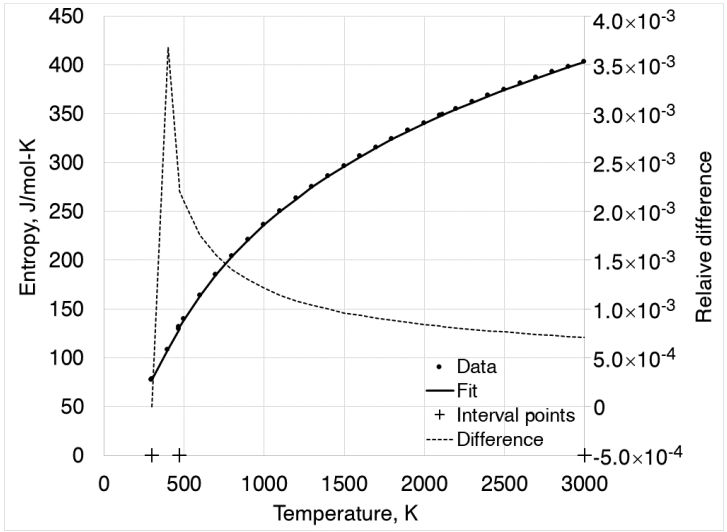


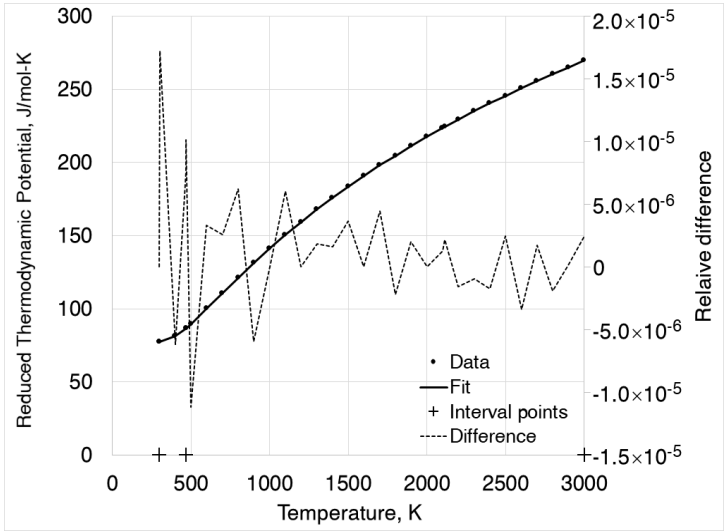
Titanium Oxide (Ti_2O_3) – O3Ti2(cr)

MW = 143.76 g/mol – $\Delta_f H^o$(298.15 K) = -1520.884 kJ/mol – S^o(298.15 K) = 77.253 J/mol-K – p^o = 0.1 MPa



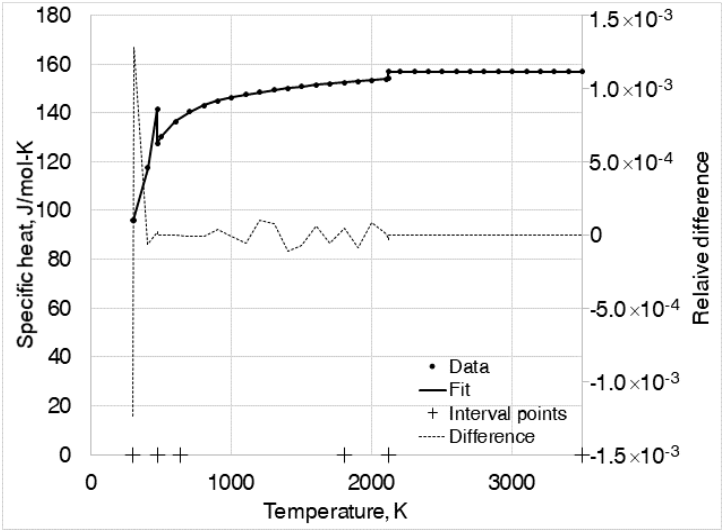


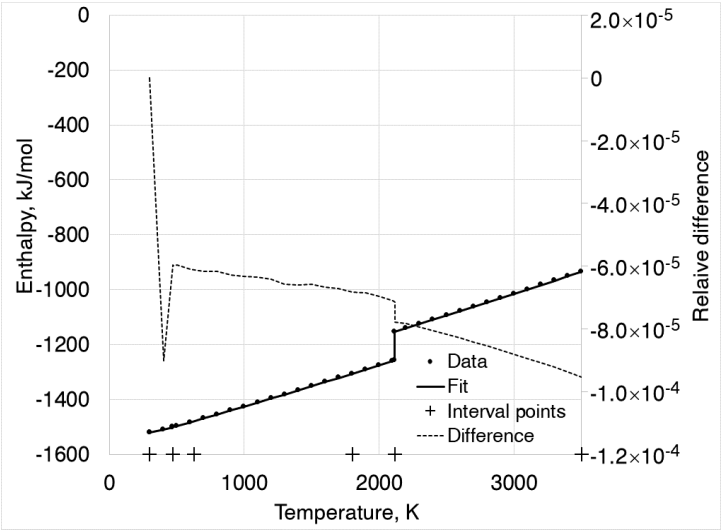


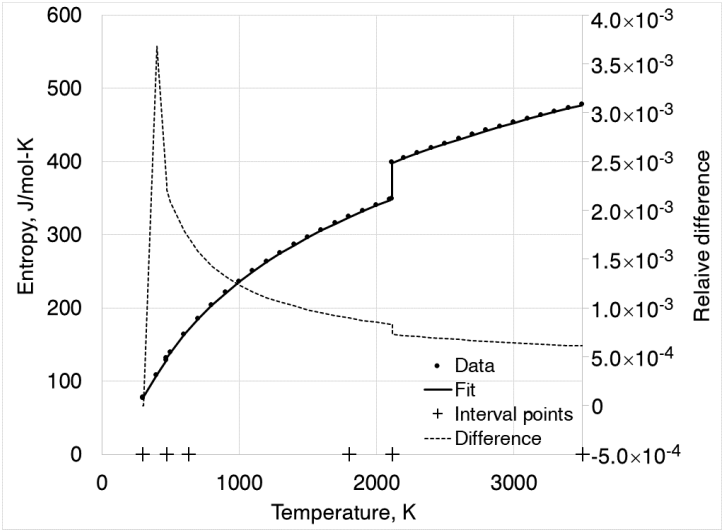


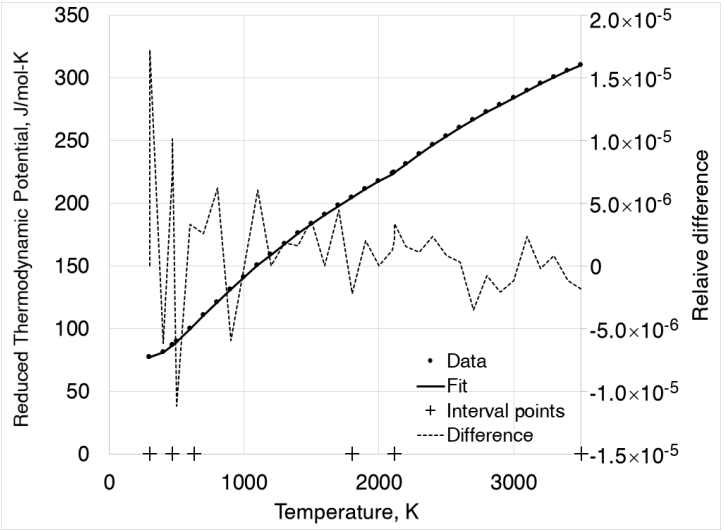
Titanium Oxide (Ti_2O_3) – O3Ti2(cr,l)

MW = 143.76 g/mol – $\Delta_f H^o$(298.15 K) = -1520.884 kJ/mol – S^o(298.15 K) = 77.253 J/mol-K – p^o = 0.1 MPa



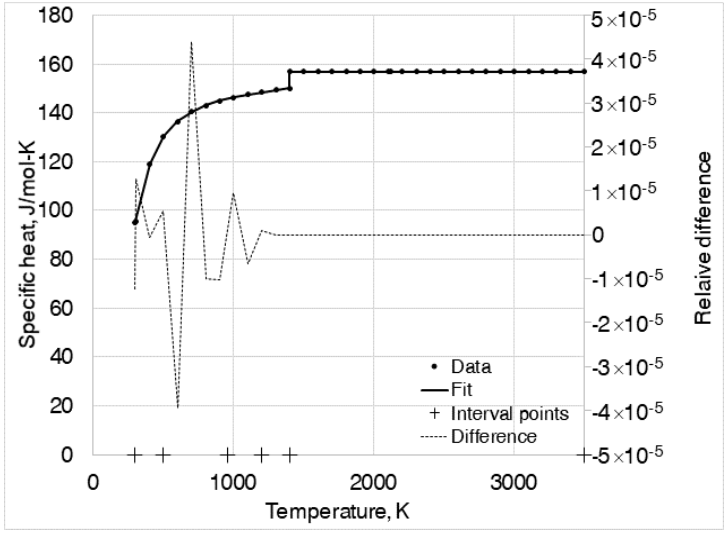


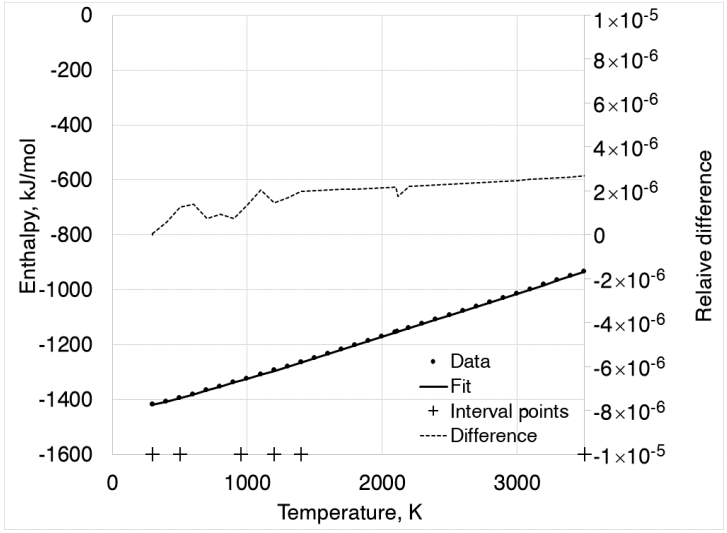


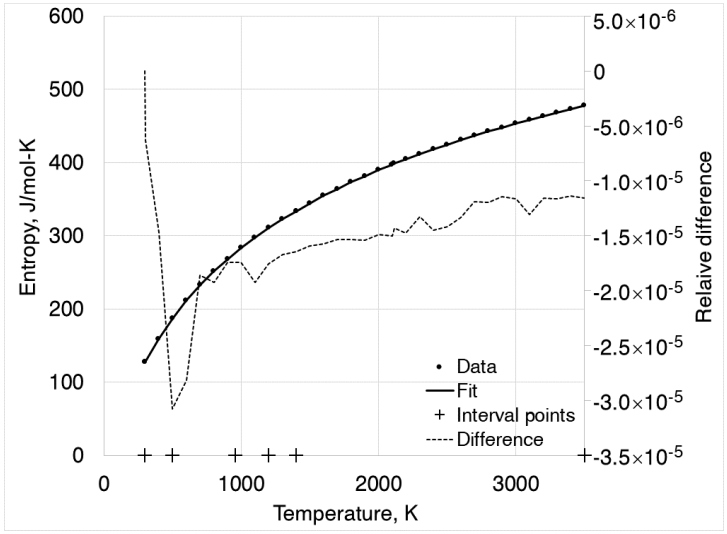


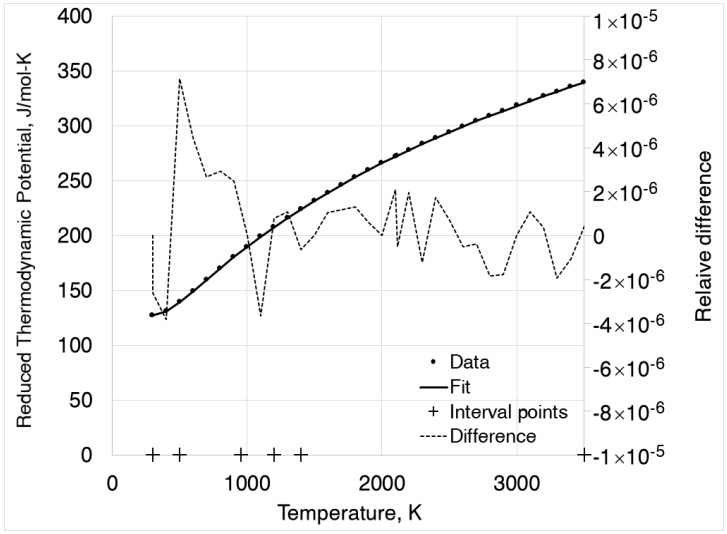
Titanium Oxide (Ti_2O_3) – O3Ti2(l)

MW = 143.76 g/mol – $\Delta_f H^o$(298.15 K) = -1418.459 kJ/mol – S^o(298.15 K) = 127.126 J/mol-K – p^o = 0.1 MPa



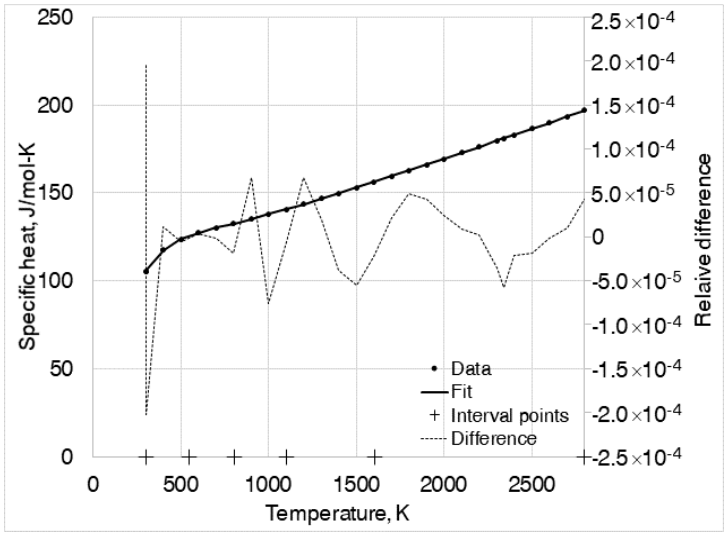


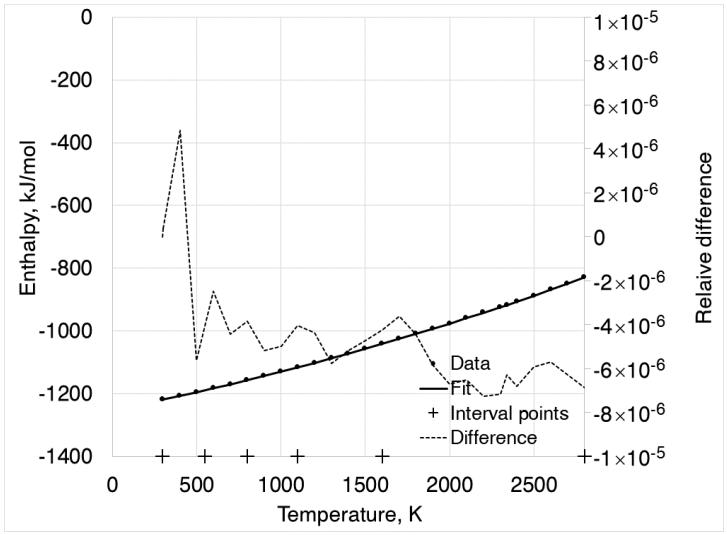


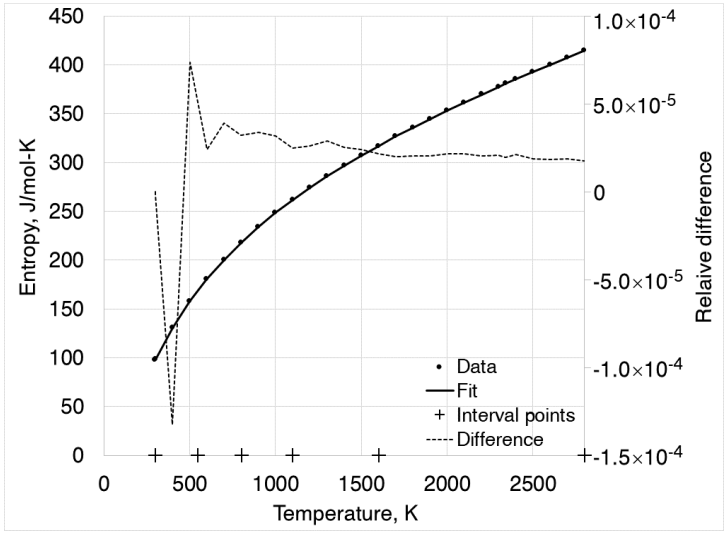


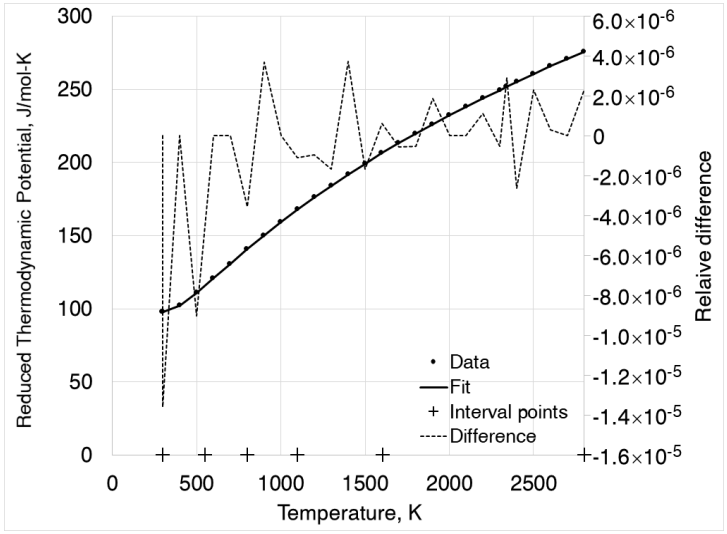
Vanadium Oxide (V_2O_3) – O3V2(cr)

MW = 149.8812 g/mol – $\Delta_f H^o$(298.15 K) = -1218.799 kJ/mol – S^o(298.15 K) = 98.067 J/mol-K – p^o = 0.1 MPa



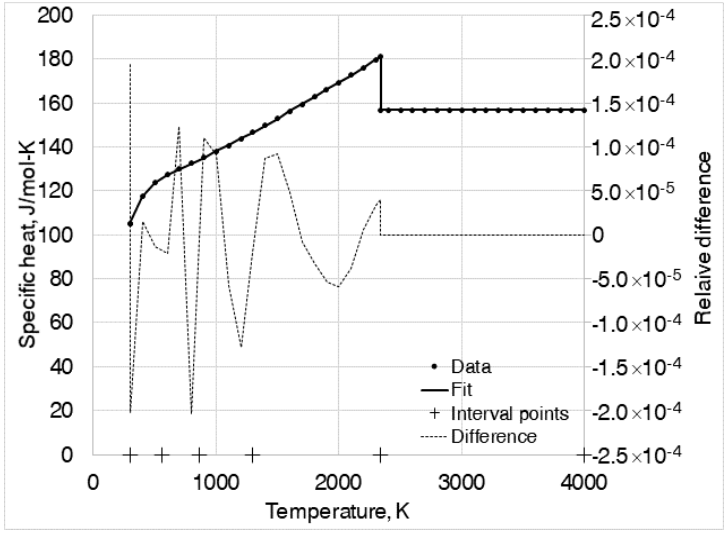


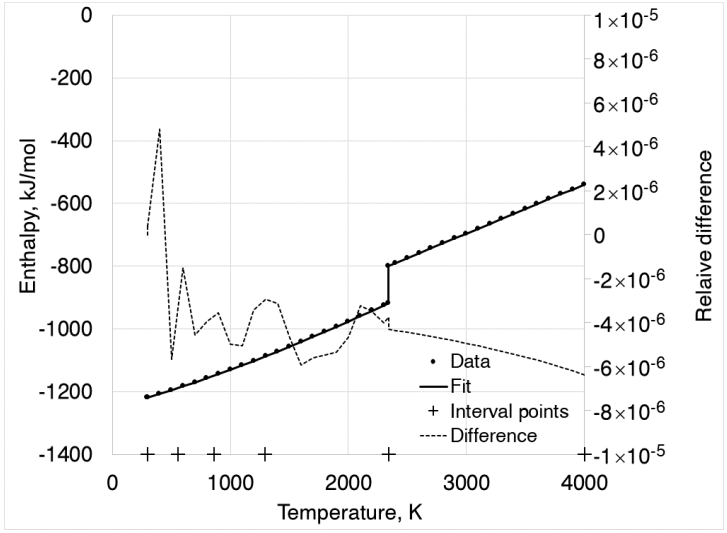


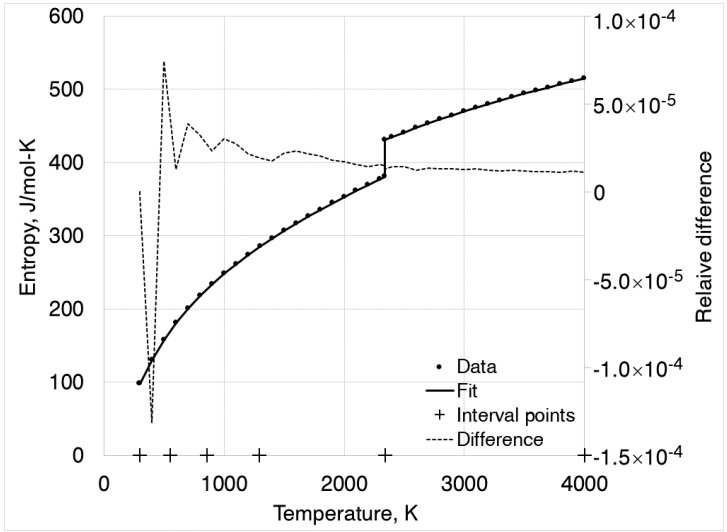


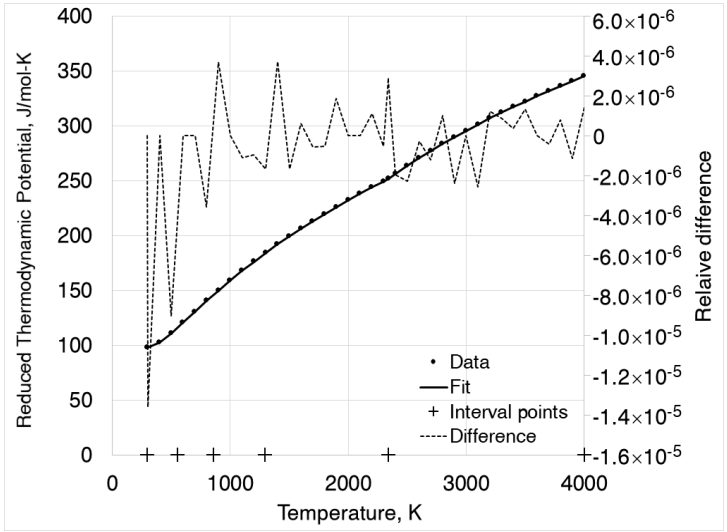
Vanadium Oxide (V_2O_3) – O3V2(cr,l)

MW = 149.8812 g/mol – $\Delta_f H^o$(298.15 K) = -1218.799 kJ/mol – S^o(298.15 K) = 98.067 J/mol-K – p^o = 0.1 MPa



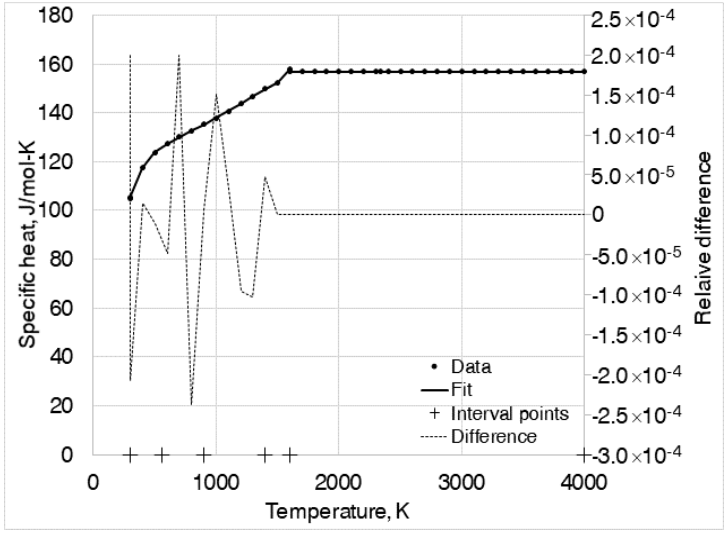


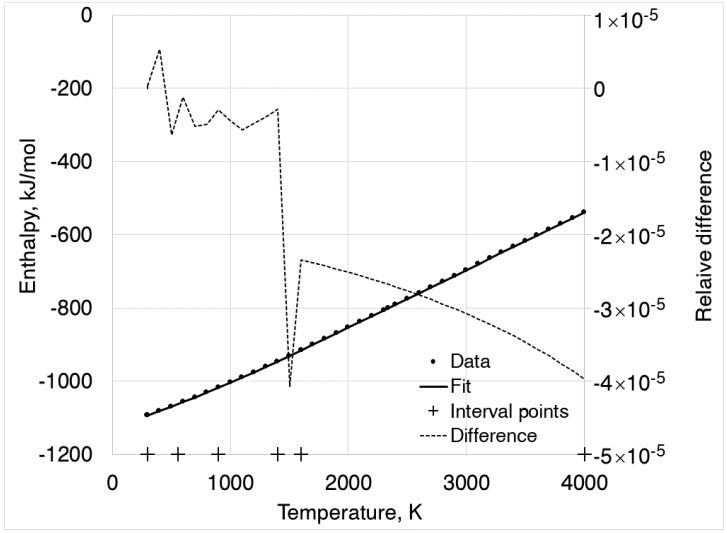


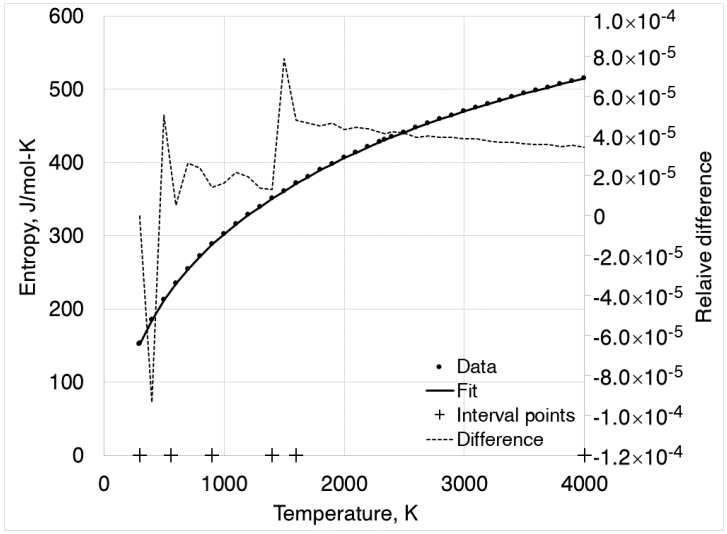


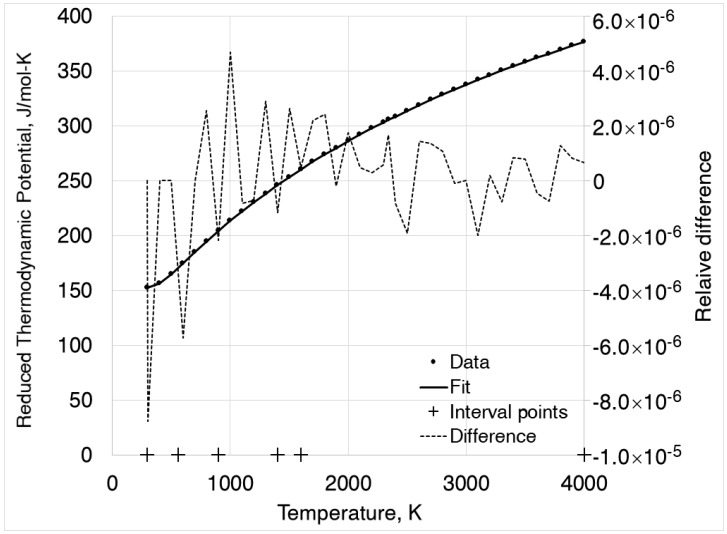
Vanadium Oxide (V_2O_3) – O3V2(l)

MW = 149.8812 g/mol – $\Delta_f H^o$(298.15 K) = -1093.054 kJ/mol – S^o(298.15 K) = 152.25 J/mol-K – p^o = 0.1 MPa



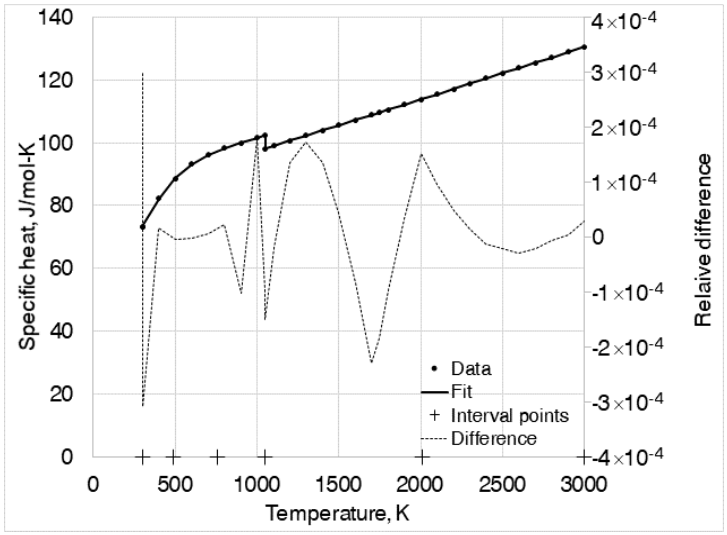


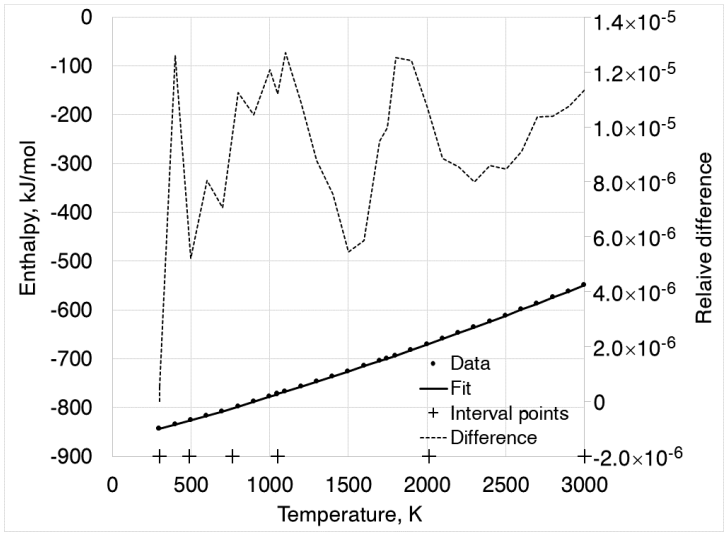


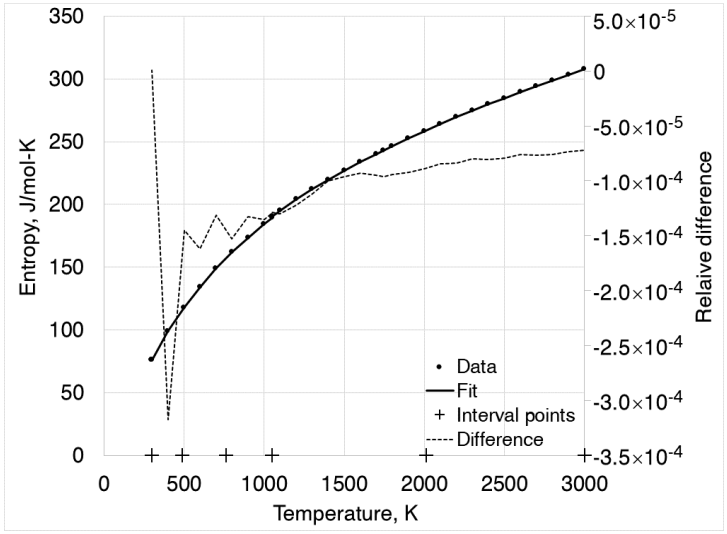


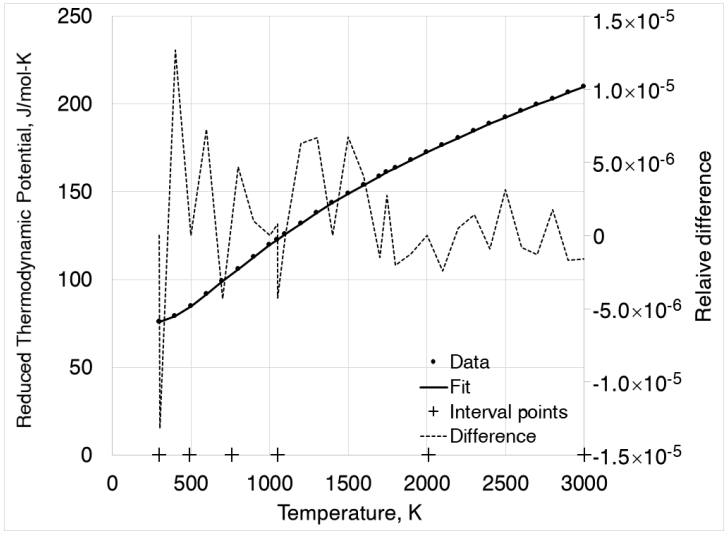
Tungsten Oxide (WO_3) – O3W1(cr)

MW = 231.85 g/mol – $\Delta_f H^o$(298.15 K) = -842.909 kJ/mol – S^o(298.15 K) = 75.912 J/mol-K – p^o = 0.1 MPa



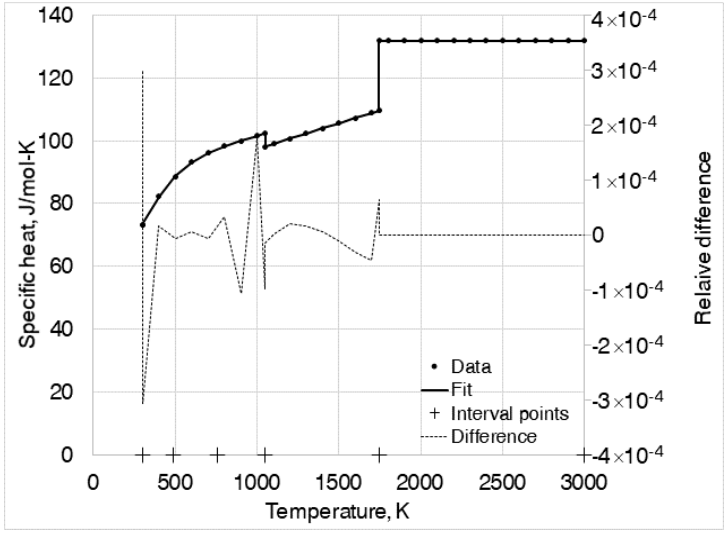


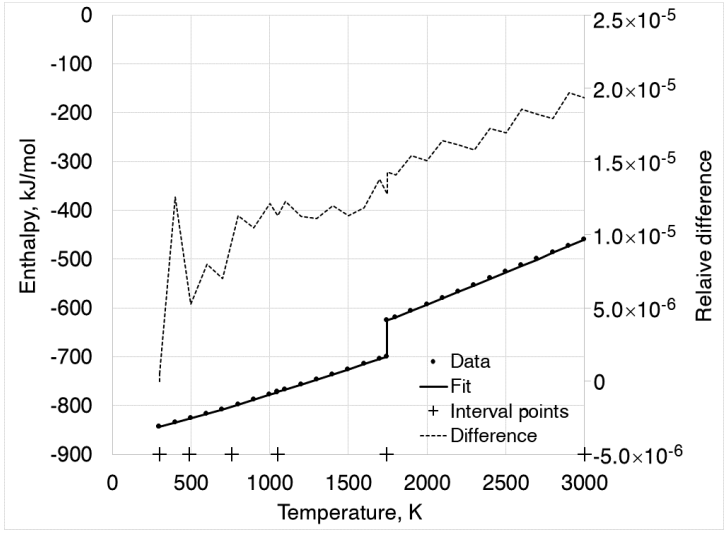


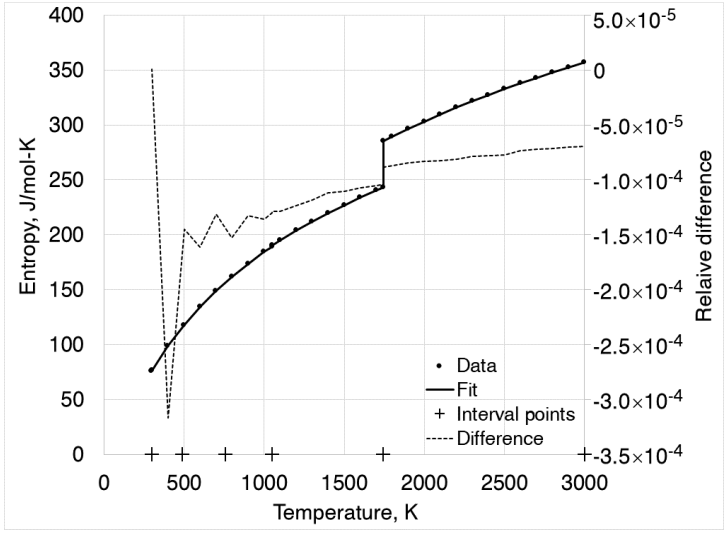


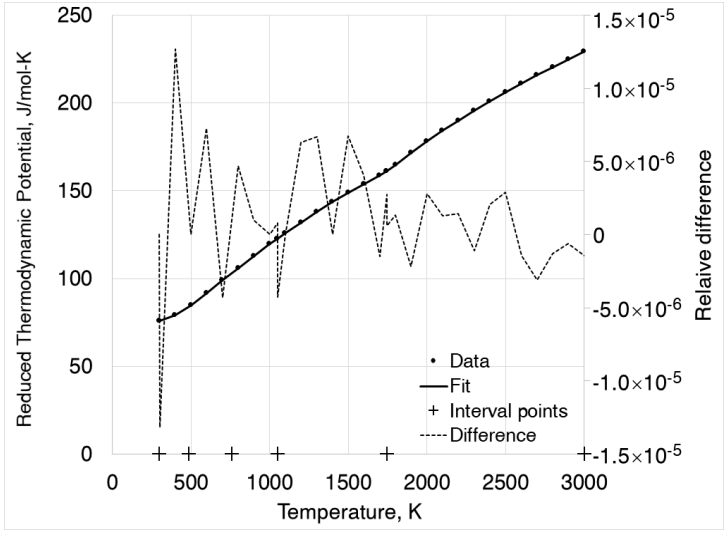
Tungsten Oxide (WO_3) – O3W1(cr,l)

MW = 231.85 g/mol – $\Delta_f H^o$(298.15 K) = -842.909 kJ/mol – S^o(298.15 K) = 75.912 J/mol-K – p^o = 0.1 MPa



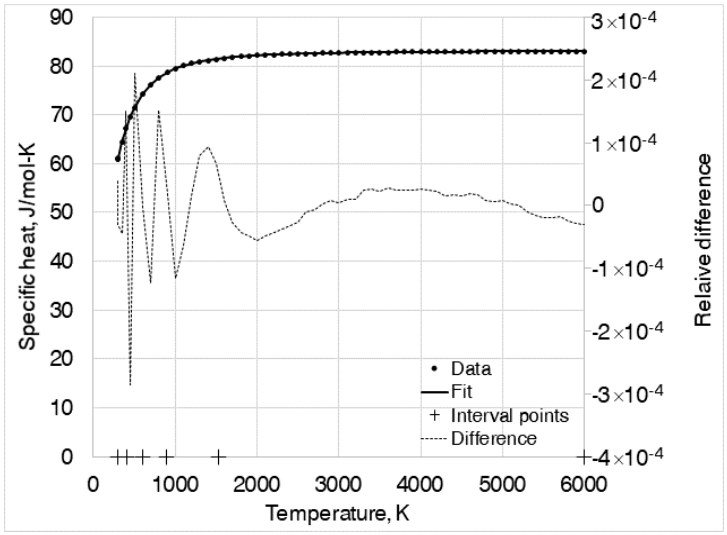


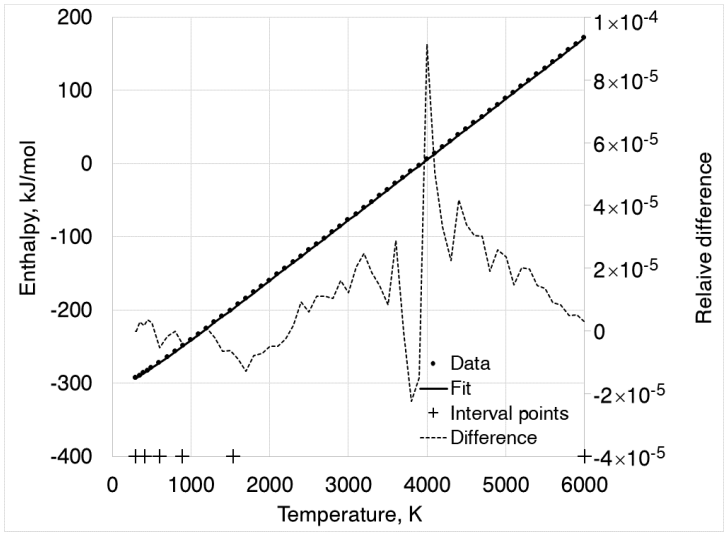


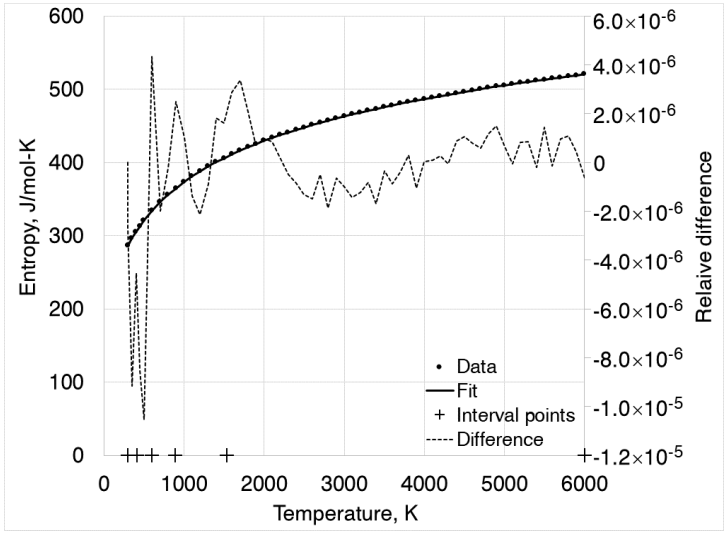


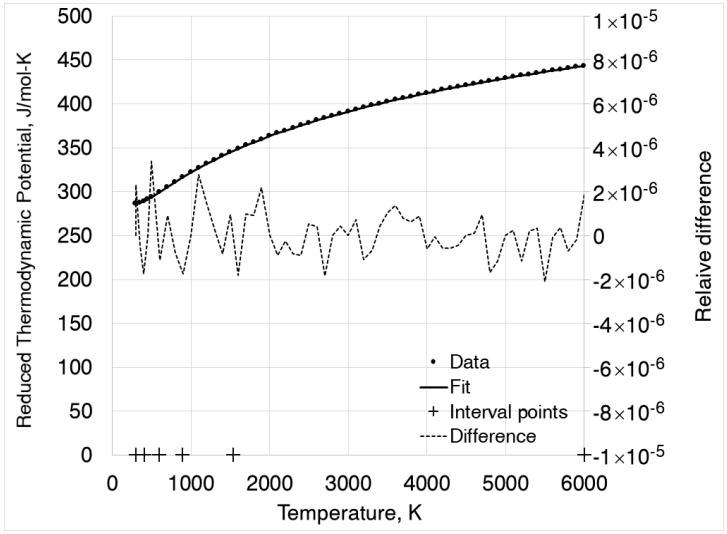
Tungsten Oxide (WO_3) – O3W1(g)

MW = 231.85 g/mol – $\Delta_f H^o$(298.15 K) = -292.88 kJ/mol – S^o(298.15 K) = 286.44 J/mol-K – p^o = 0.1 MPa



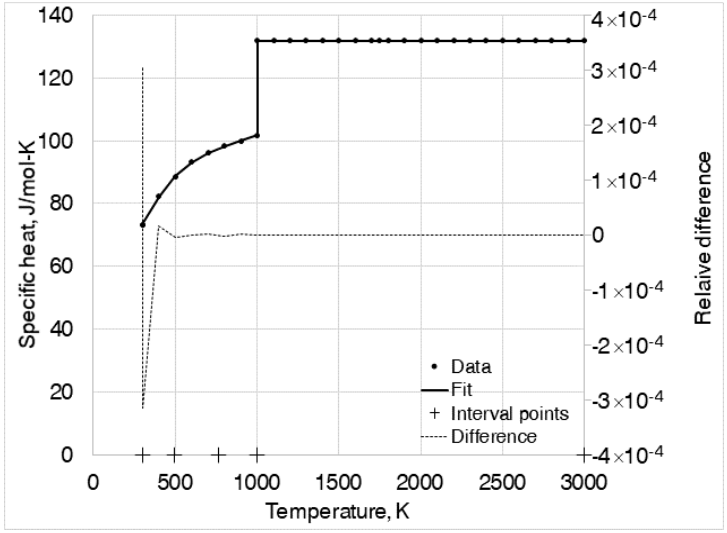


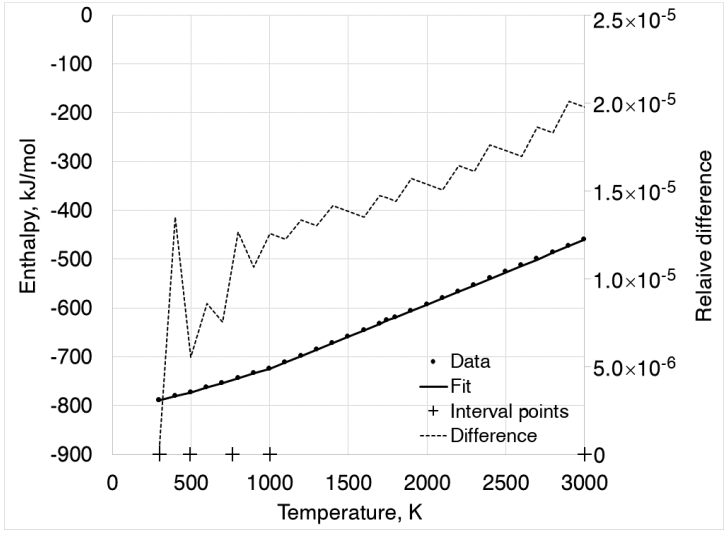


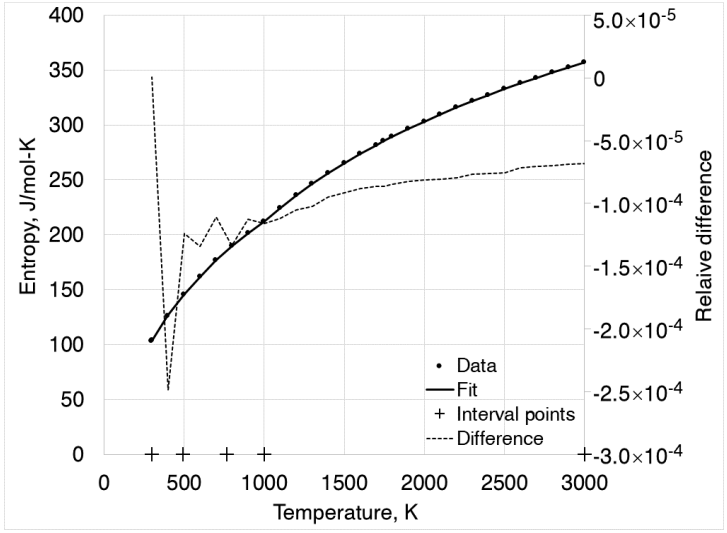


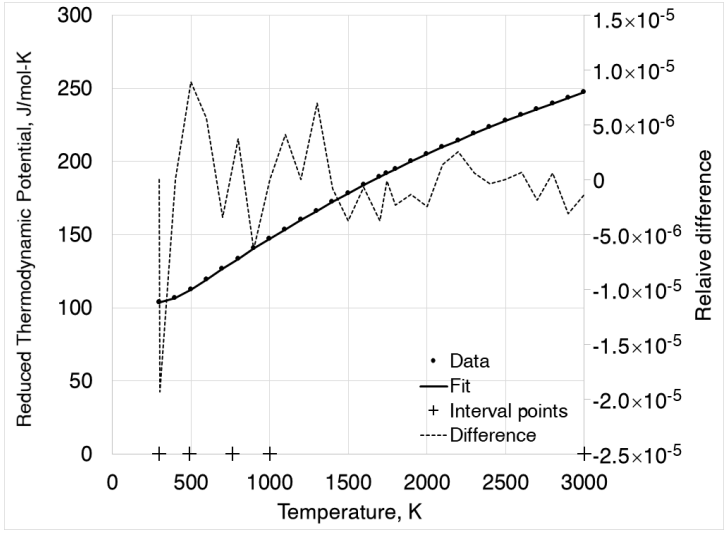
Tungsten Oxide (WO_3) – O3W1(l)

MW = 231.85 g/mol – $\Delta_f H^o$(298.15 K) = -788.915 kJ/mol – S^o(298.15 K) = 103.508 J/mol-K – p^o = 0.1 MPa



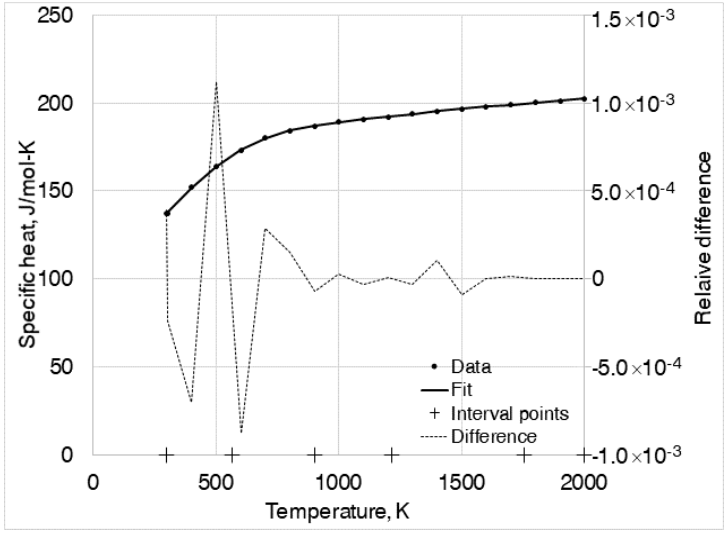


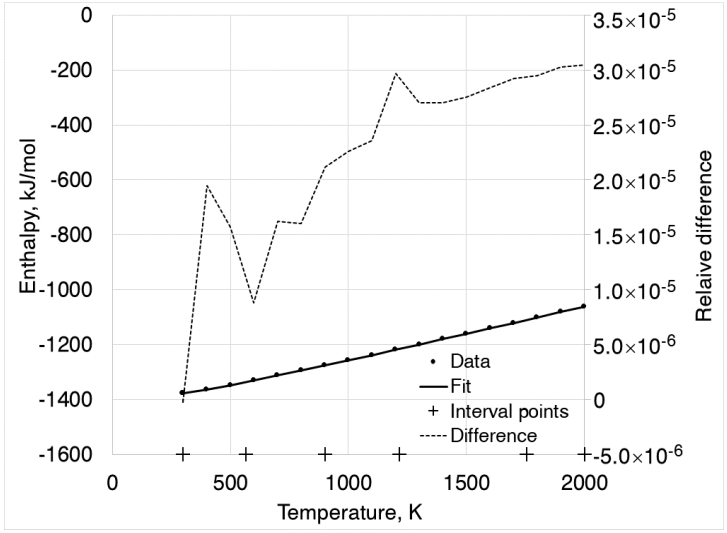


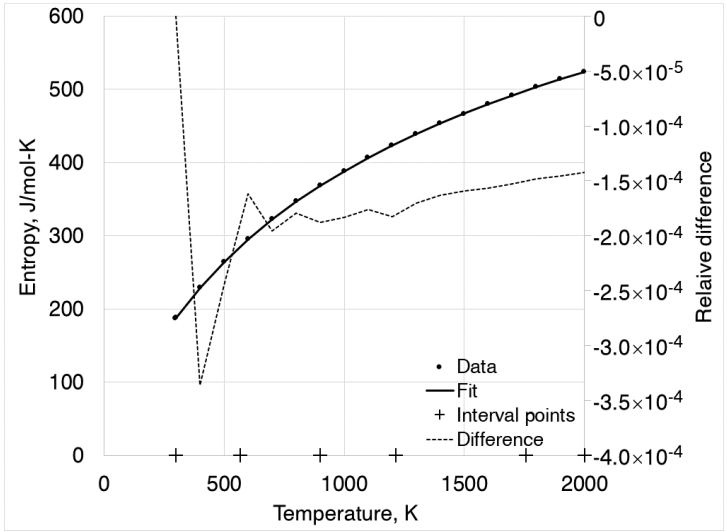


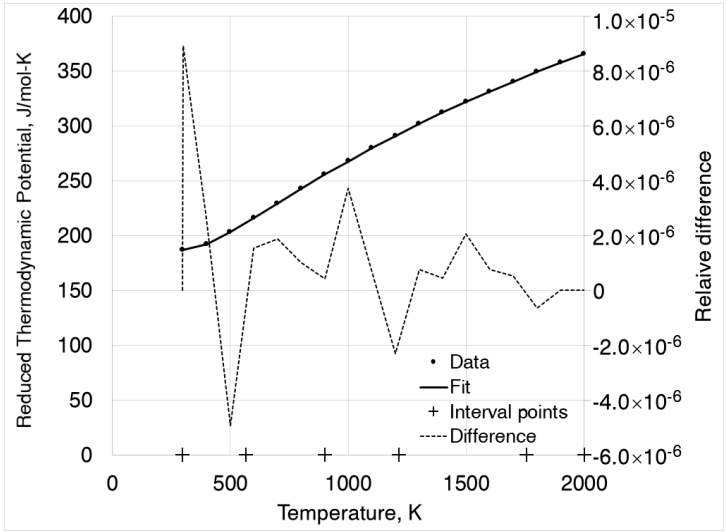
Lead Silicate (Pb_2SiO_4) – O4Pb2Si1(cr)

MW = 506.5 g/mol – $\Delta_f H^o$(298.15 K) = -1377.875 kJ/mol – S^o(298.15 K) = 186.832 J/mol-K – p^o = 0.1 MPa



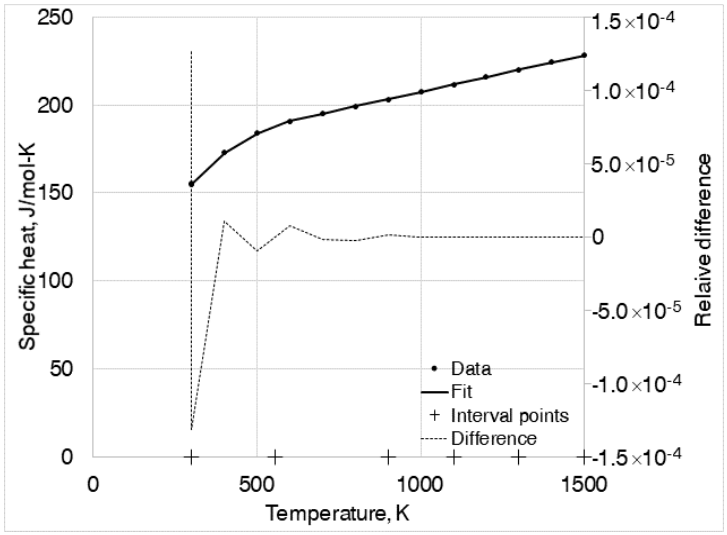


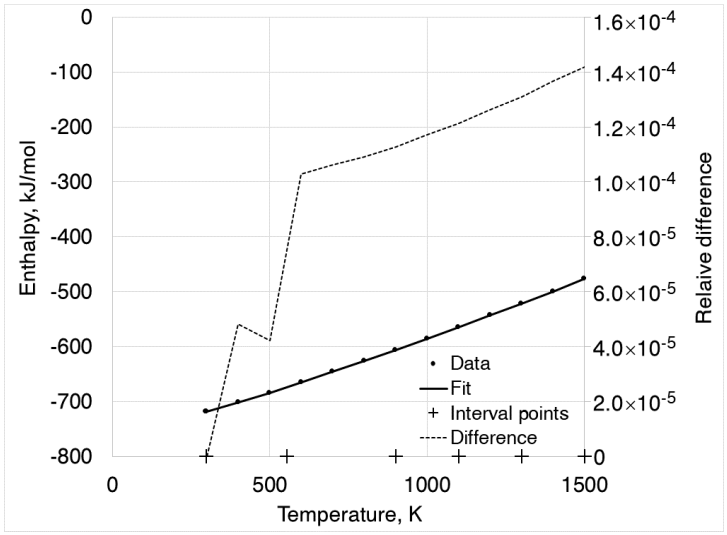


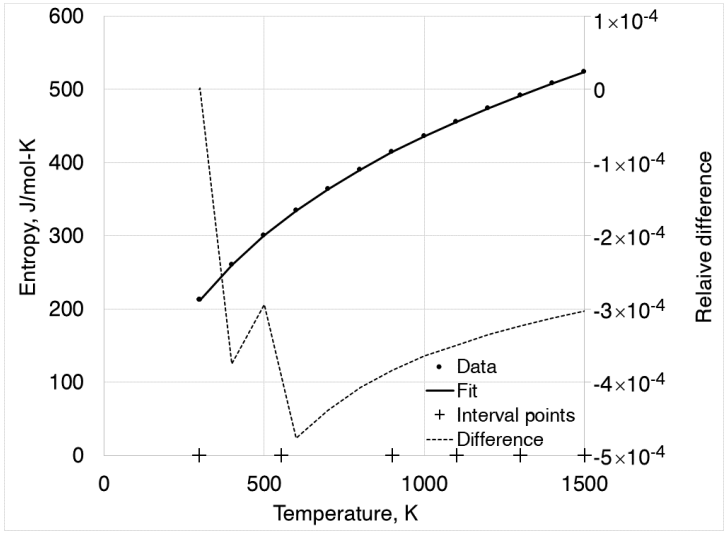


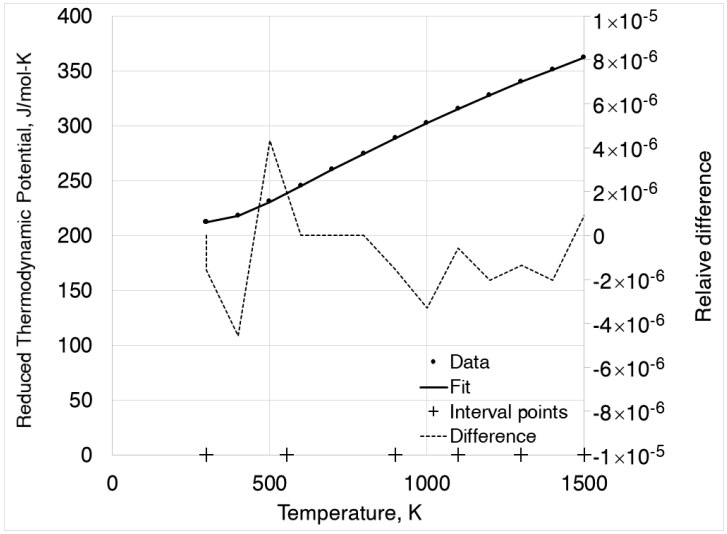
Lead Oxide (Pb_3O_4) – O4Pb3(cr)

MW = 685.6 g/mol – $\Delta_f H^o$(298.15 K) = -718.686 kJ/mol – S^o(298.15 K) = 211.961 J/mol-K – p^o = 0.1 MPa



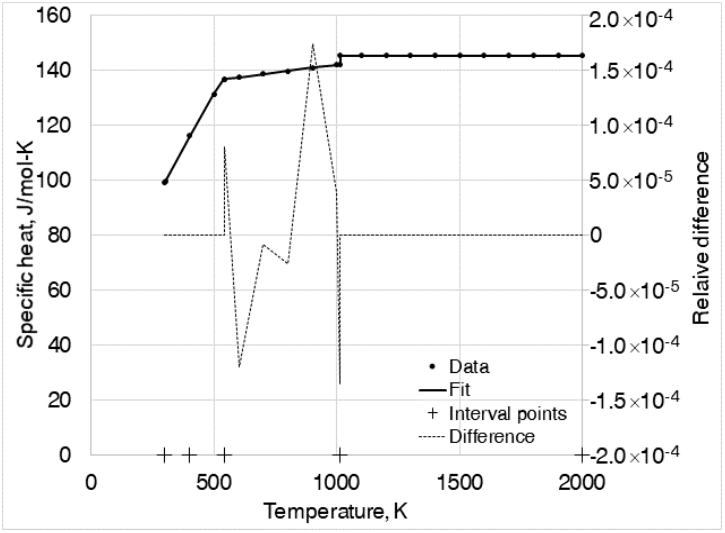


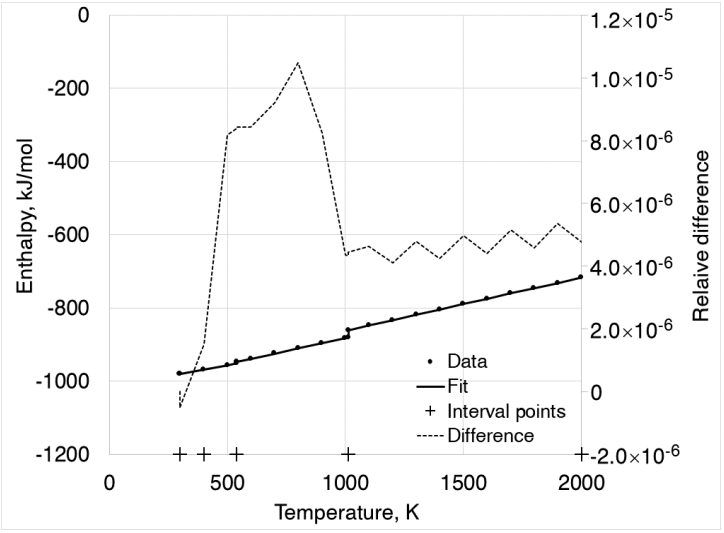


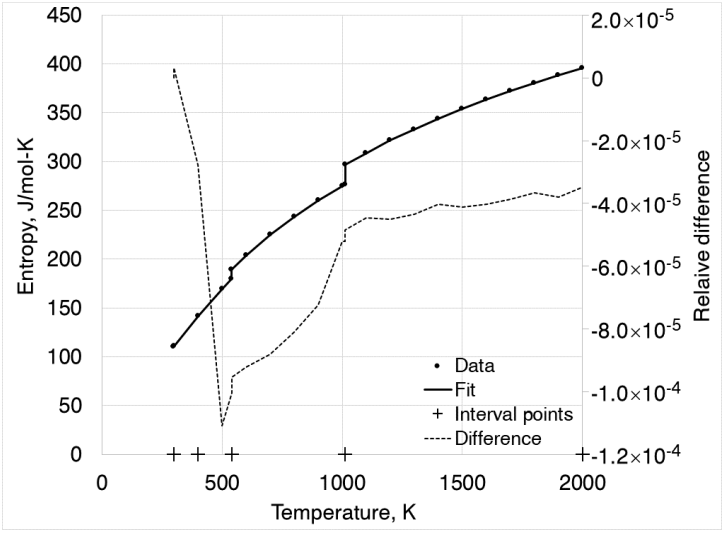


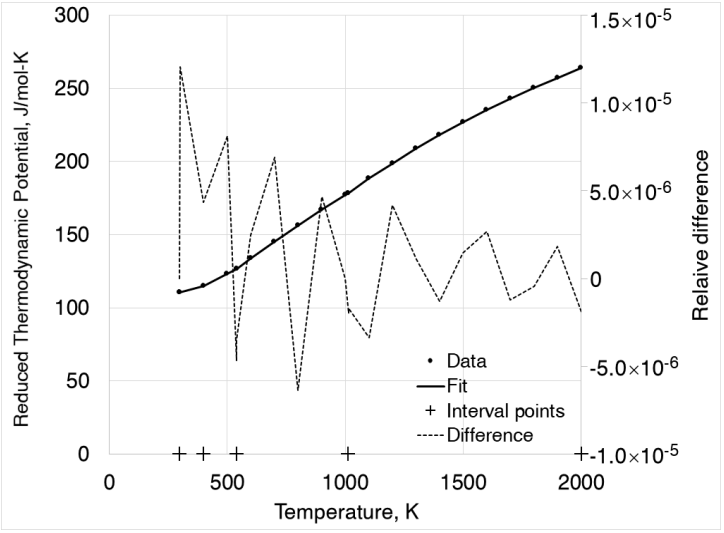
Zinc Sulfate ($ZnSO_4$) – O4S1Zn1(cr)

MW = 161.44 g/mol – $\Delta_f H^o$(298.15 K) = -980.144 kJ/mol – S^o(298.15 K) = 110.541 J/mol-K – p^o = 0.1 MPa



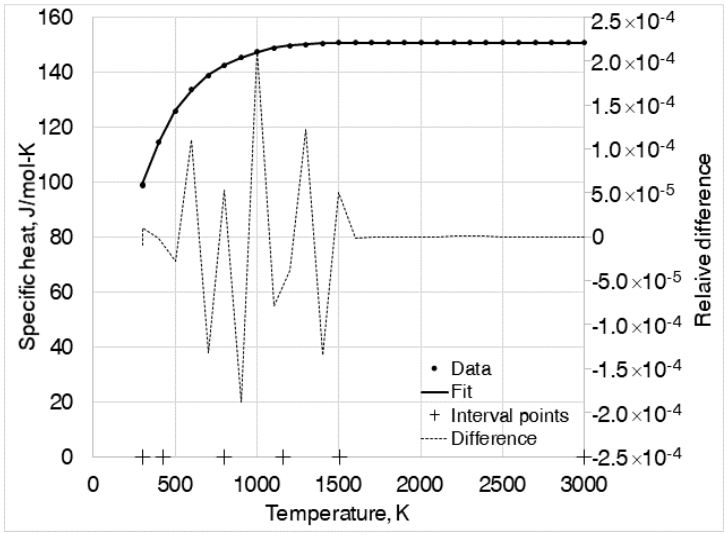


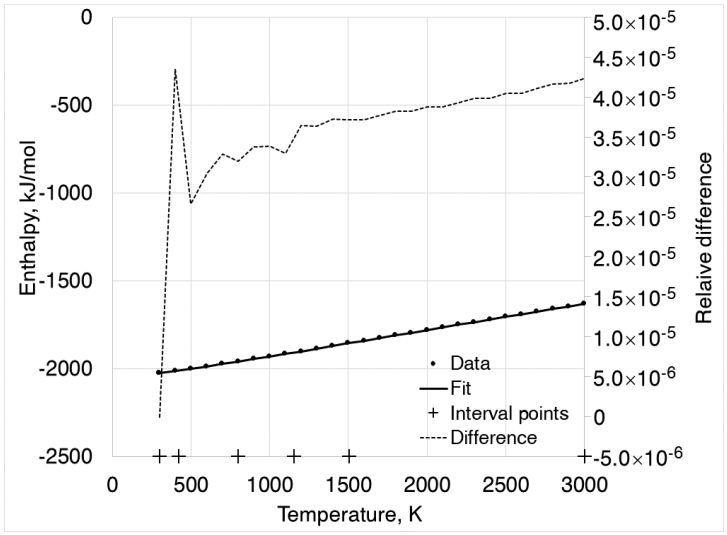


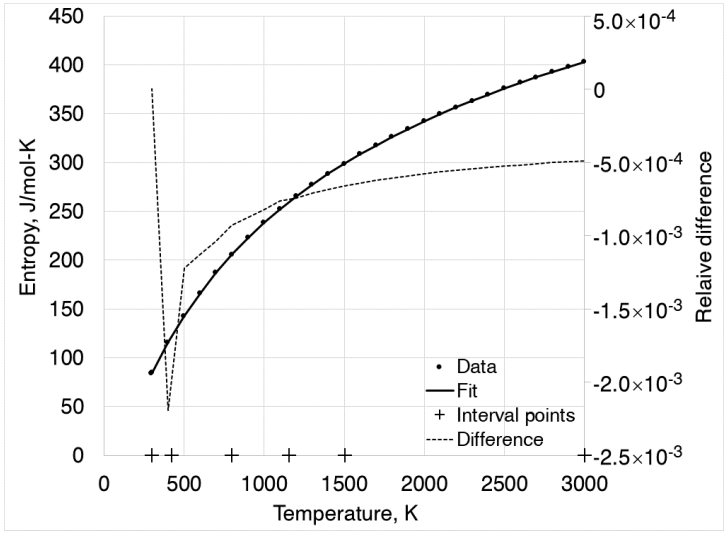


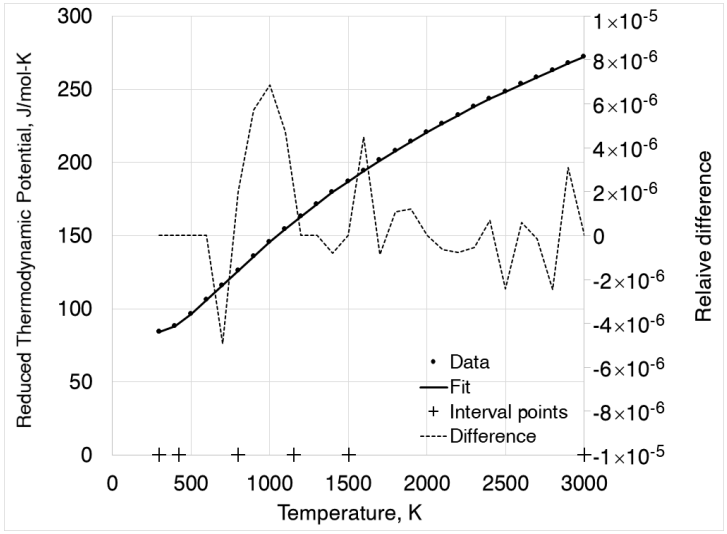
Zirconium Silicate ($ZrSiO_4$) – O4Si1Zr1(cr)

MW = 183.3 g/mol – $\Delta_f H^o$(298.15 K) = -2023.951 kJ/mol – S^o(298.15 K) = 84.026 J/mol-K – p^o = 0.1 MPa



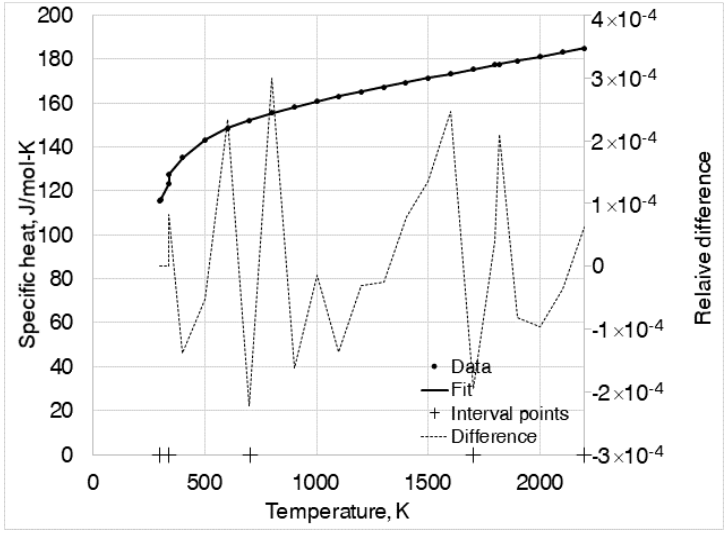


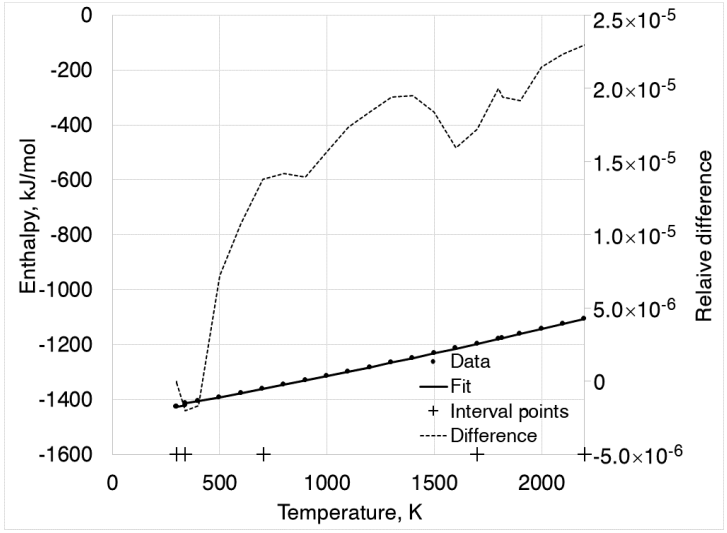


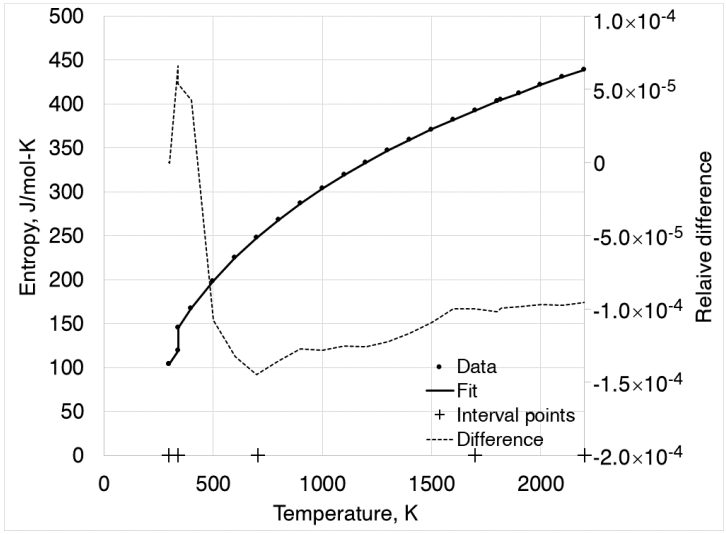


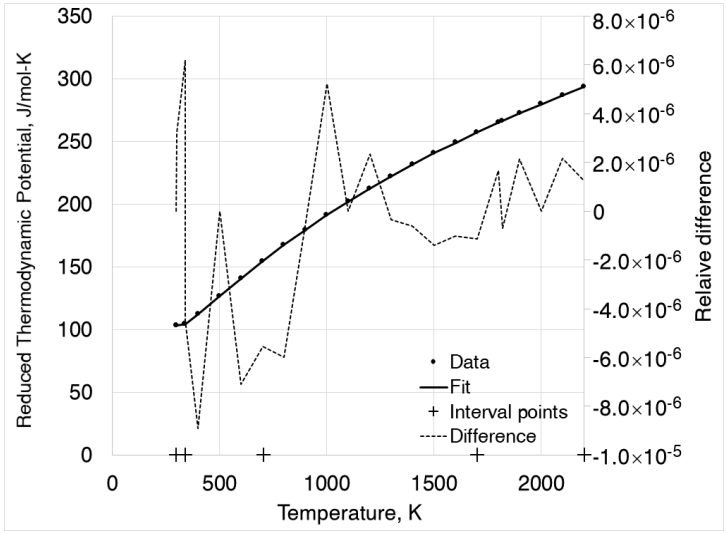
Vanadium Oxide (V_2O_4) – O4V2(cr)

MW = 165.8806 g/mol – $\Delta_f H^o$(298.15 K) = -1427.162 kJ/mol – S^o(298.15 K) = 103.525 J/mol-K – p^o = 0.1 MPa



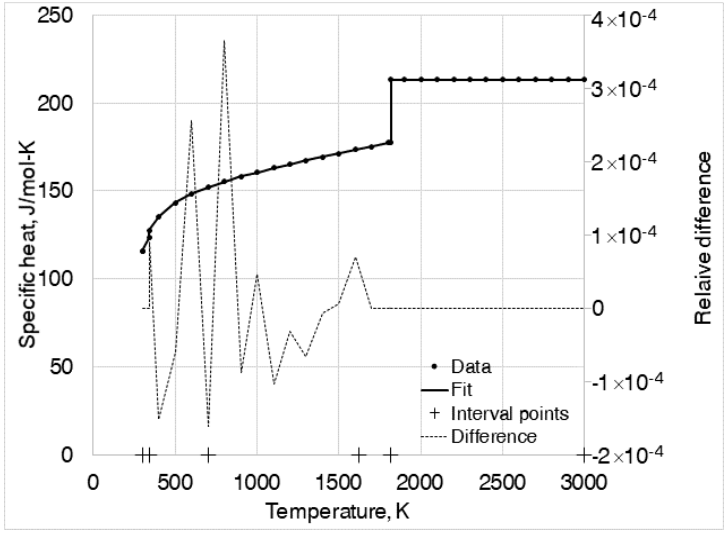


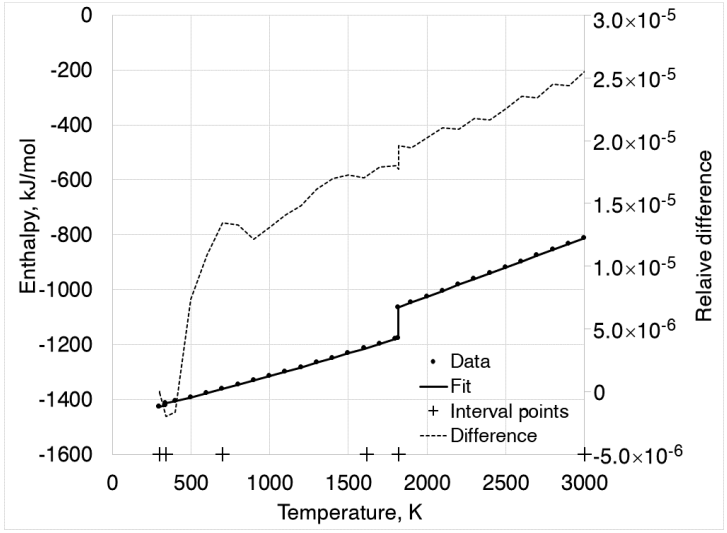


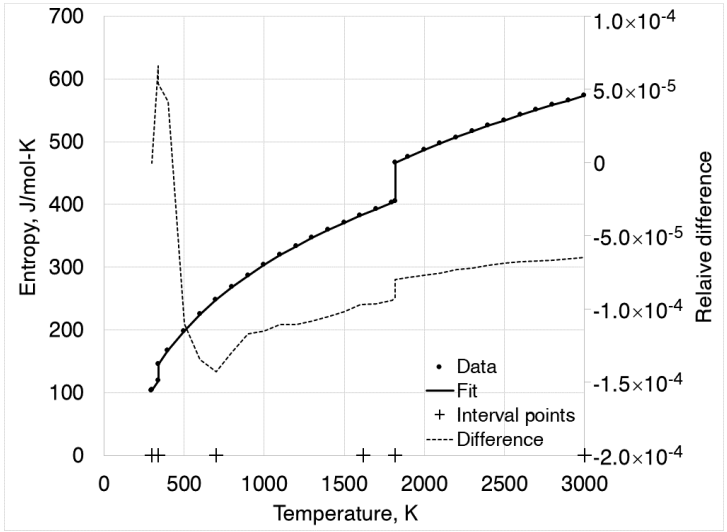


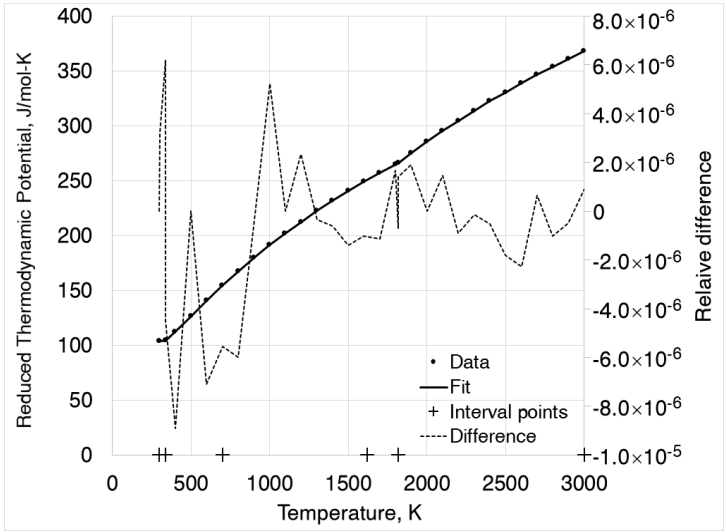
Vanadium Oxide (V_2O_4) – O4V2(cr,l)

MW = 165.8806 g/mol – $\Delta_f H^o$(298.15 K) = -1427.162 kJ/mol – S^o(298.15 K) = 103.525 J/mol-K – p^o = 0.1 MPa



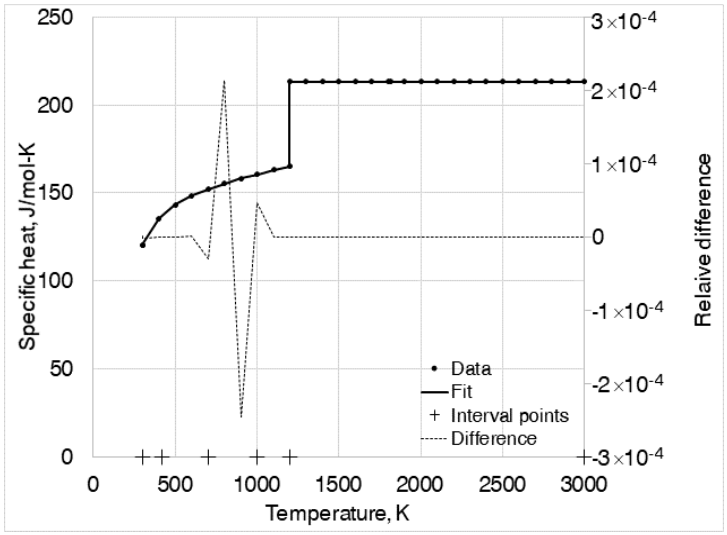


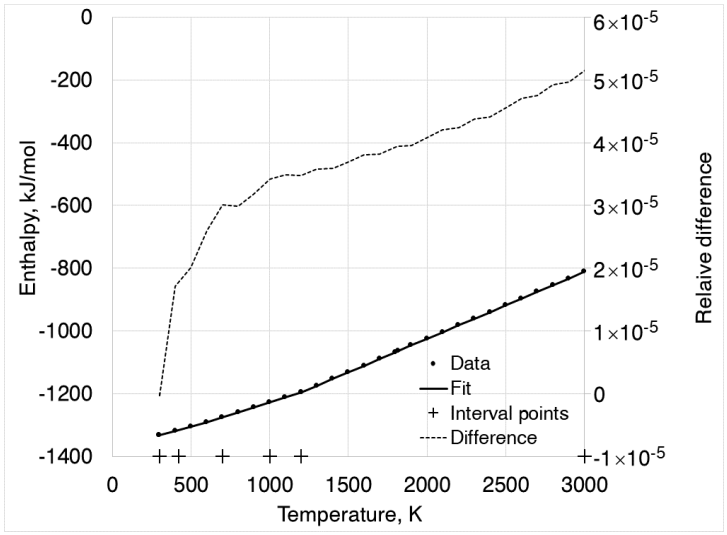


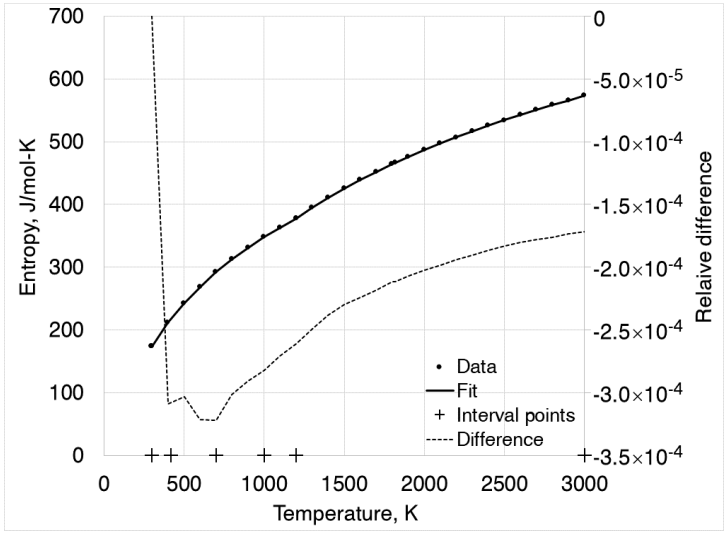


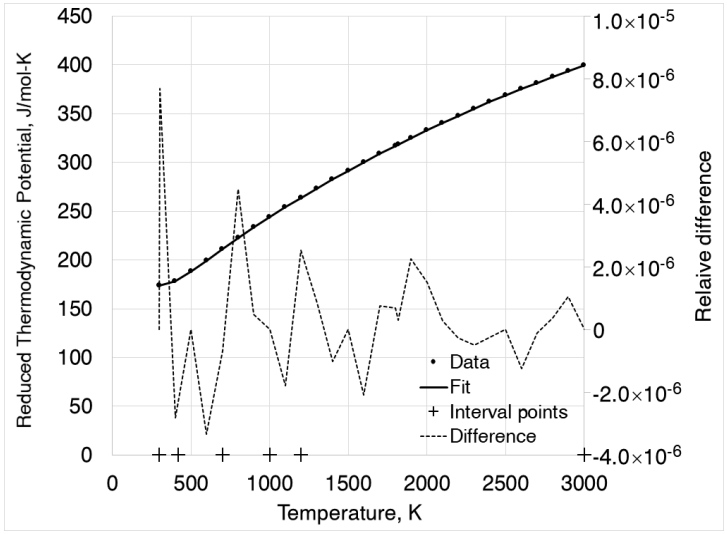
Vanadium Oxide (V_2O_4) – O4V2(l)

MW = 165.8806 g/mol – $\Delta_f H^o$(298.15 K) = -1332.151 kJ/mol – S^o(298.15 K) = 173.473 J/mol-K – p^o = 0.1 MPa



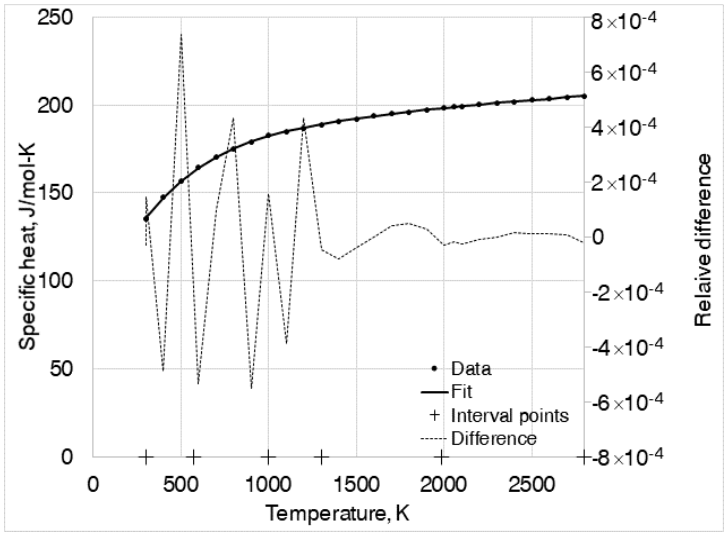


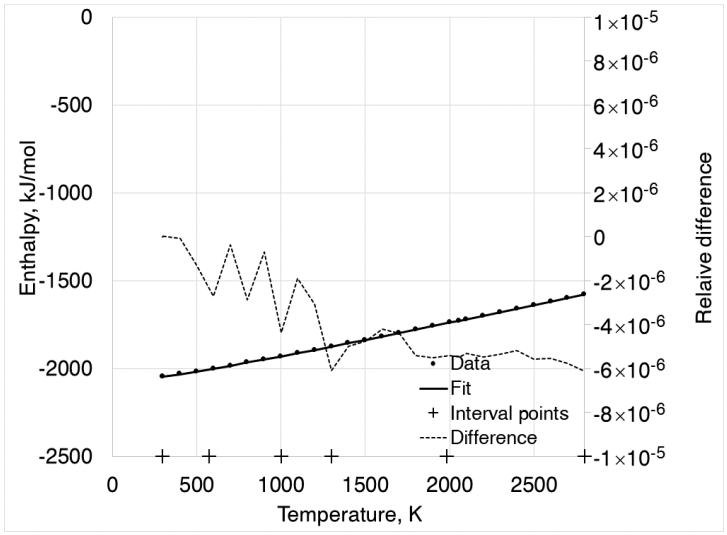


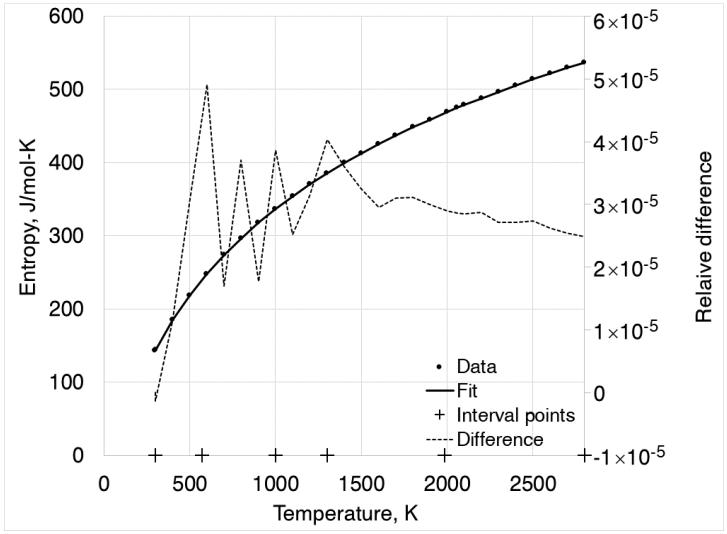


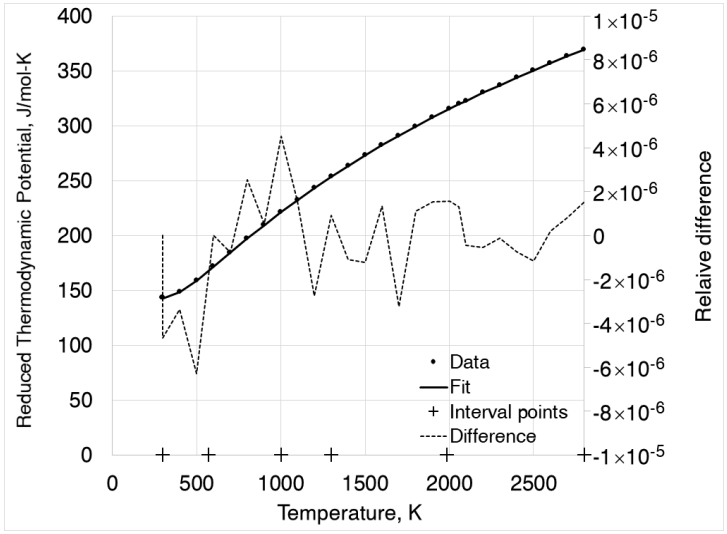
Tantalum Oxide (Ta_2O_5) – O5Ta2(cr)

MW = 441.8928 g/mol – $\Delta_f H^o$(298.15 K) = -2045.976 kJ/mol – S^o(298.15 K) = 143.114 J/mol-K – p^o = 0.1 MPa



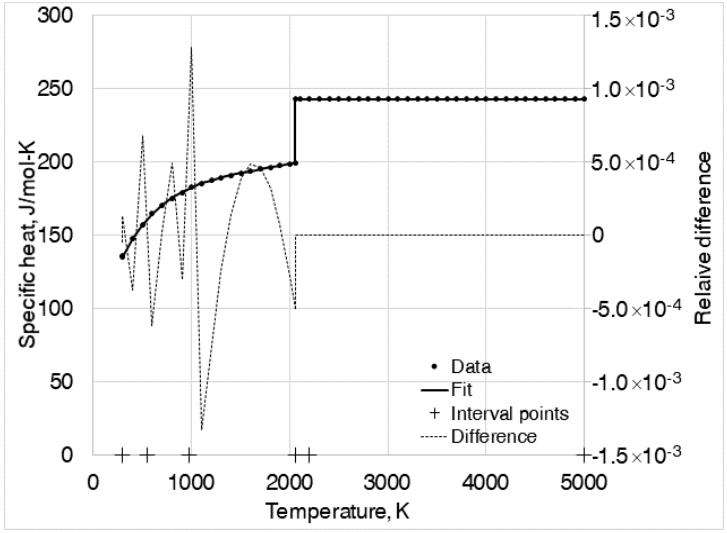


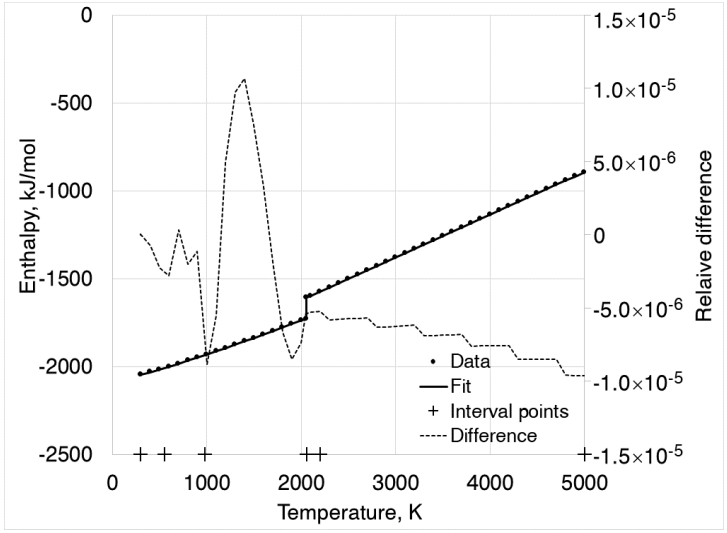


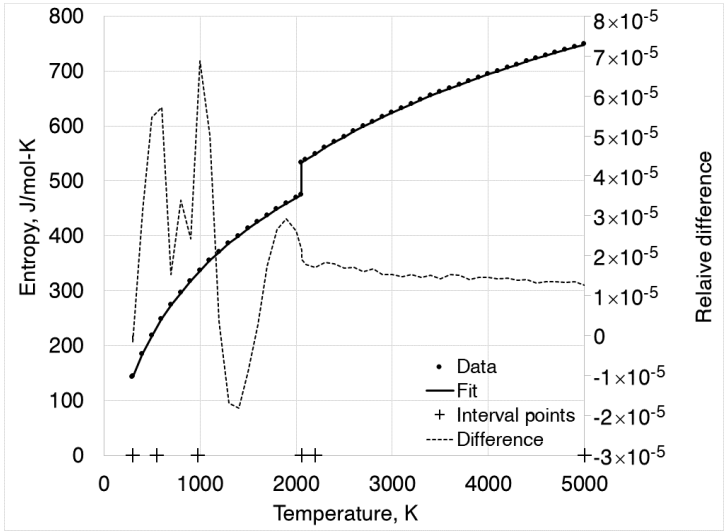


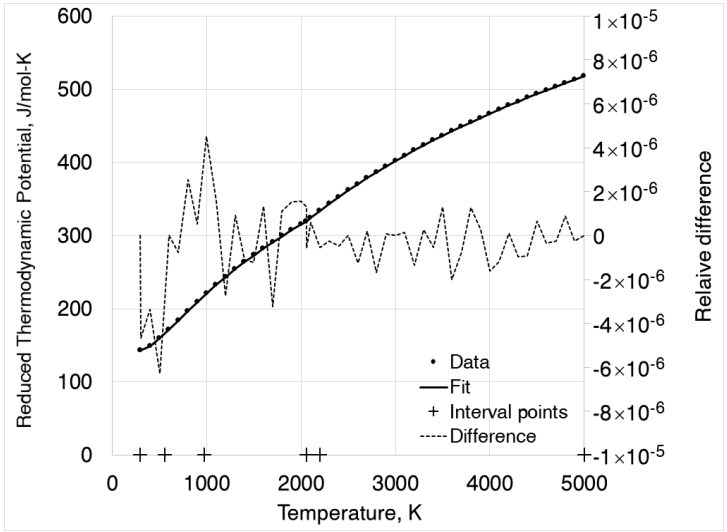
Tantalum Oxide (Ta_2O_5) – O5Ta2(cr,l)

MW = 441.8928 g/mol – $\Delta_f H^o$(298.15 K) = -2045.976 kJ/mol – S^o(298.15 K) = 143.114 J/mol-K – p^o = 0.1 MPa



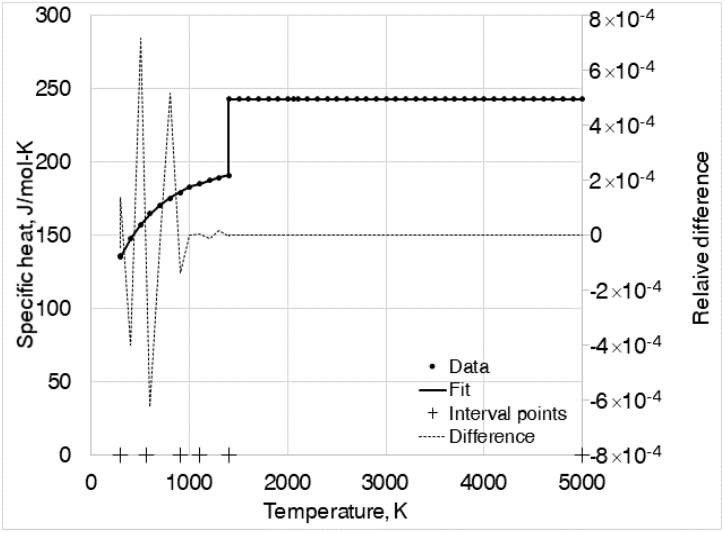


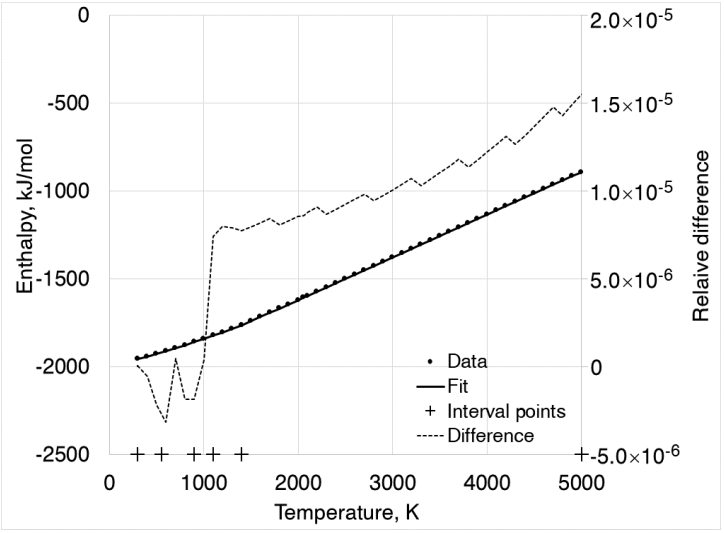


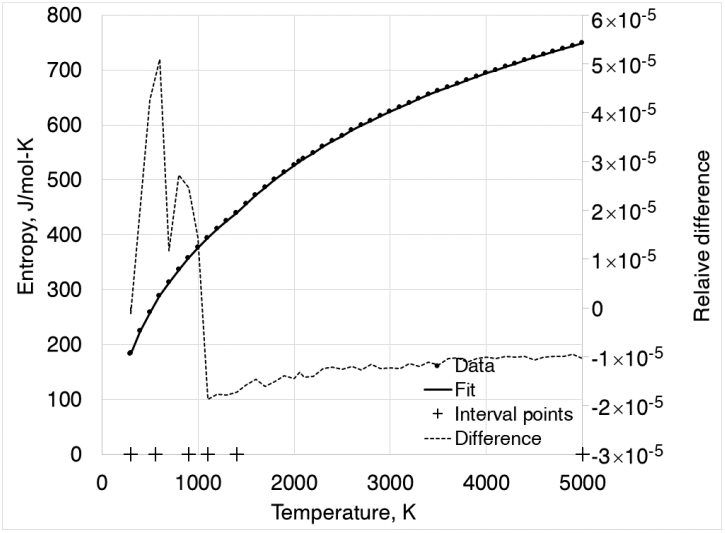


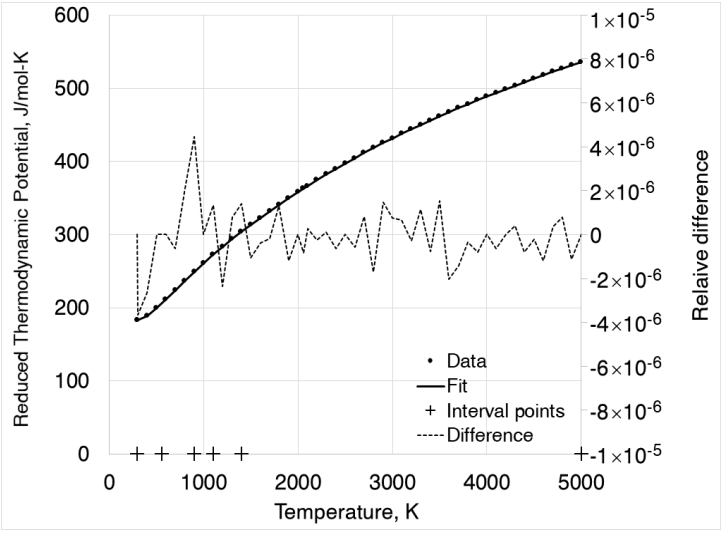
Tantalum Oxide (Ta_2O_5) – O5Ta2(l)

MW = 441.8928 g/mol – $\Delta_f H^o$(298.15 K) = -1957.171 kJ/mol – S^o(298.15 K) = 183.05 J/mol-K – p^o = 0.1 MPa



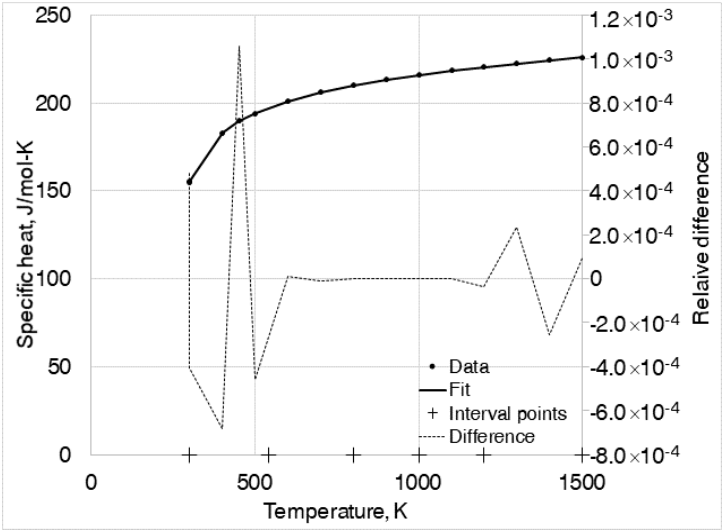


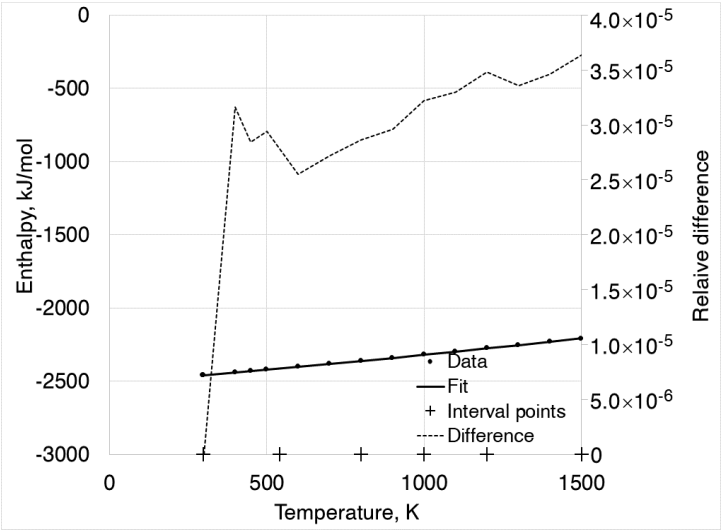


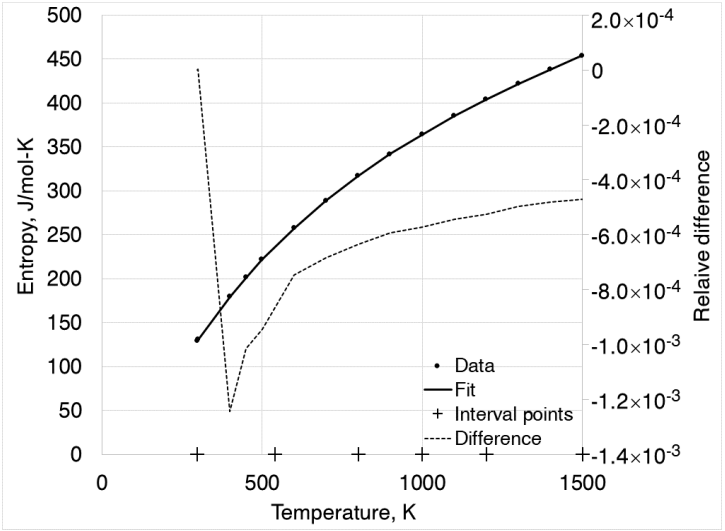


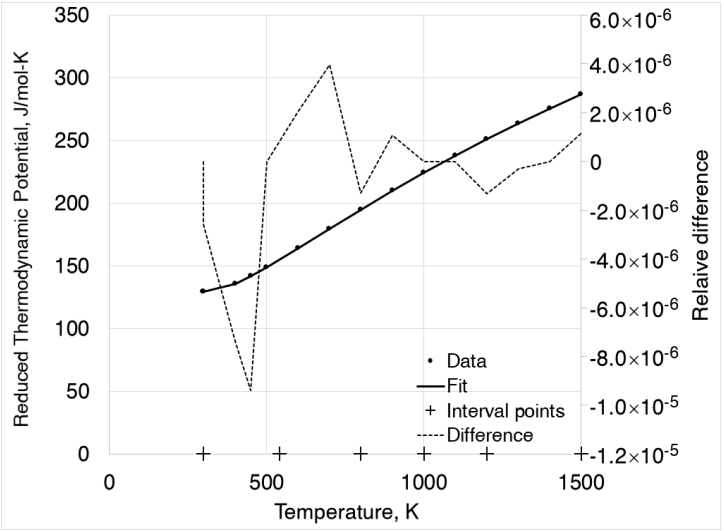
Titanium Oxide, Alpha (Ti_3O_5) – O5Ti3(cr)_01

MW = 223.64 g/mol – $\Delta_f H^o$(298.15 K) = -2459.146 kJ/mol – S^o(298.15 K) = 129.369 J/mol-K – p^o = 0.1 MPa



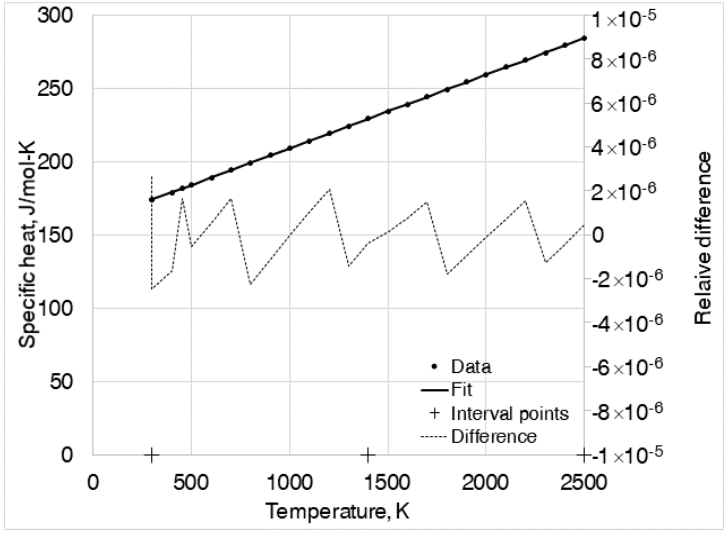


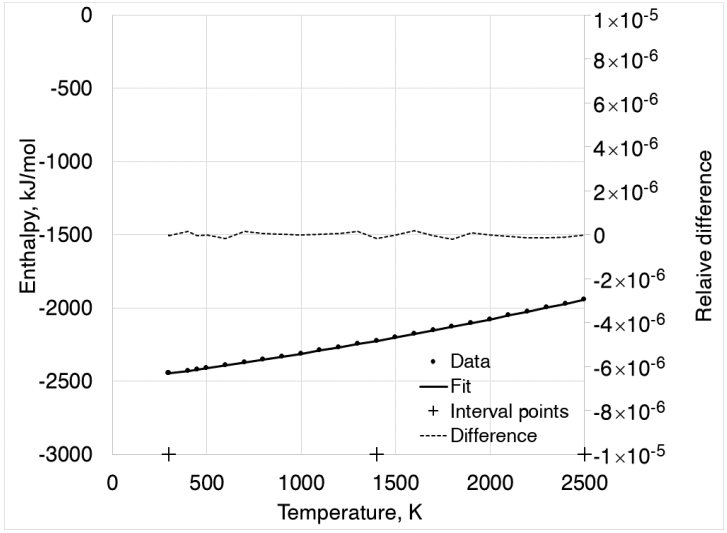


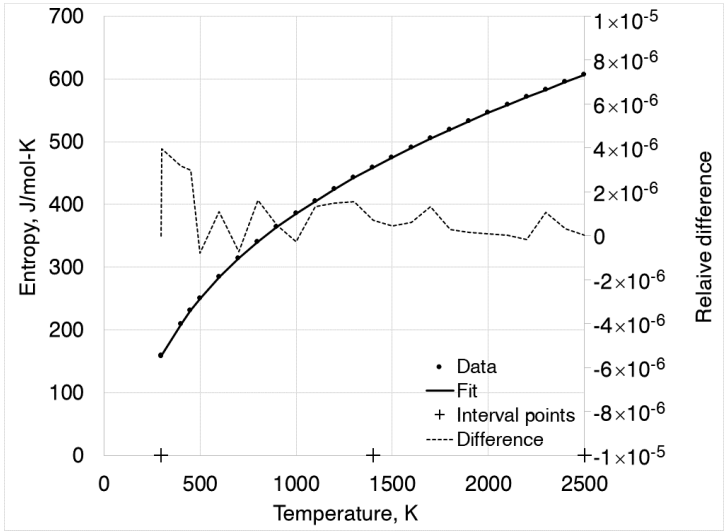


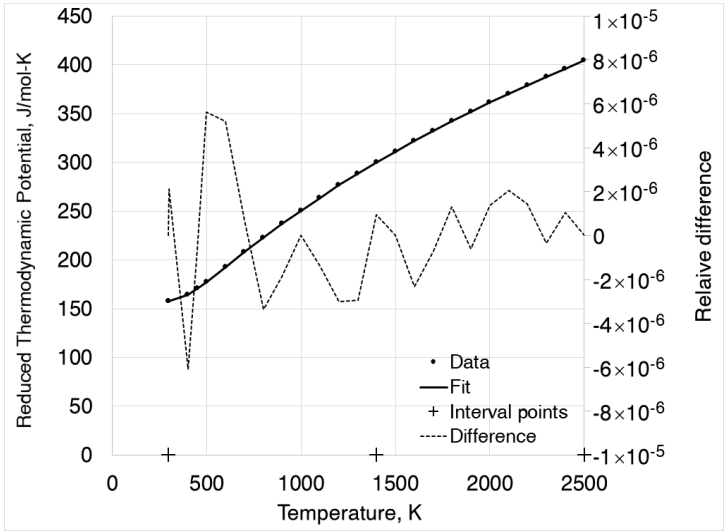
Titanium Oxide, Beta (Ti_3O_5) – O5Ti3(cr)_02

MW = 223.64 g/mol – $\Delta_f H^o$(298.15 K) = -2446.189 kJ/mol – S^o(298.15 K) = 157.617 J/mol-K – p^o = 0.1 MPa



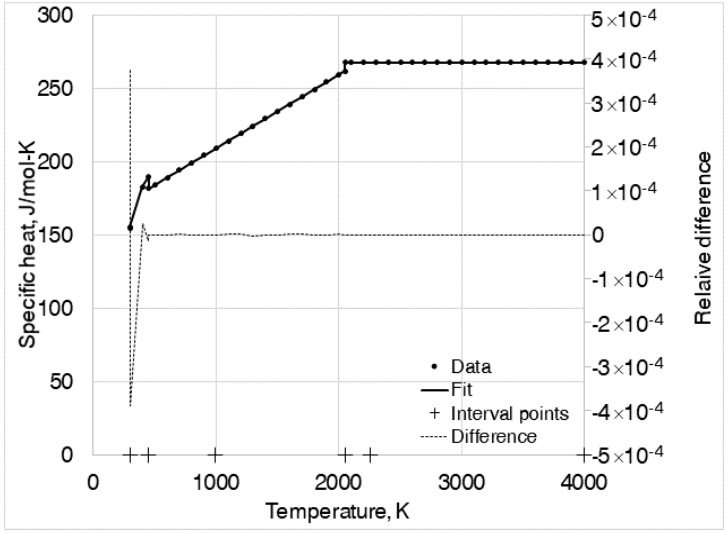


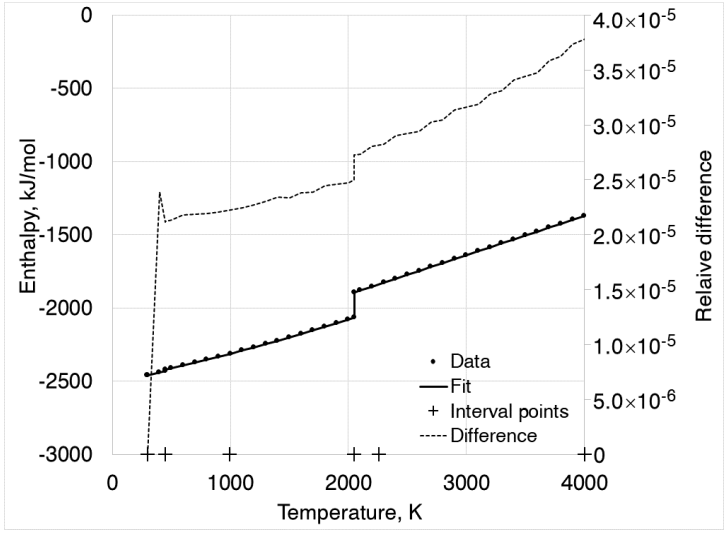


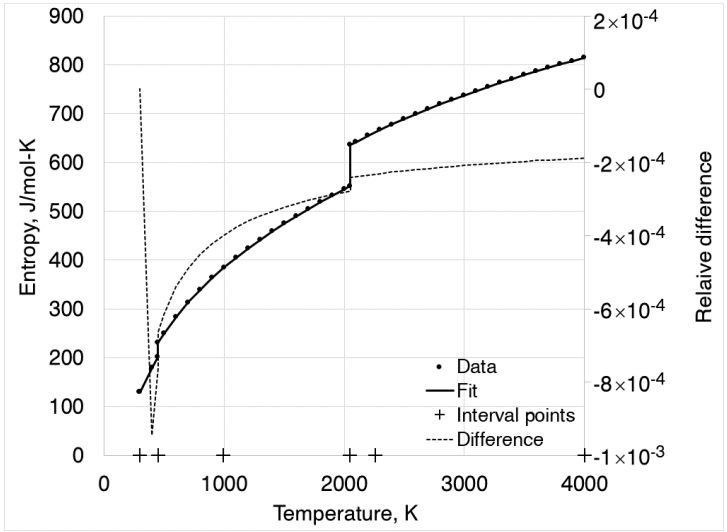


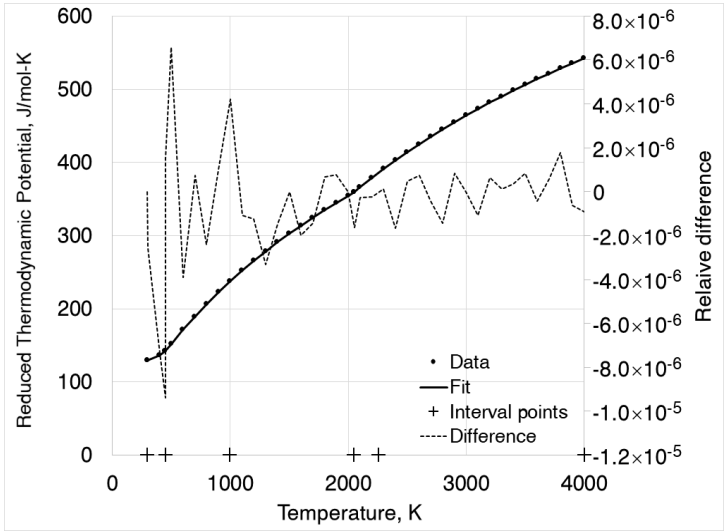
Titanium Oxide (Ti_3O_5) – O5Ti3(cr,l)

MW = 223.64 g/mol – $\Delta_f H^o$(298.15 K) = -2459.146 kJ/mol – S^o(298.15 K) = 129.369 J/mol-K – p^o = 0.1 MPa



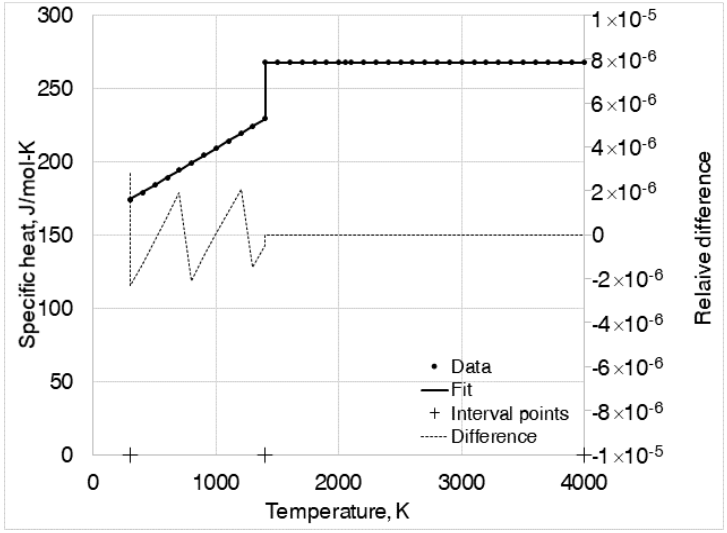


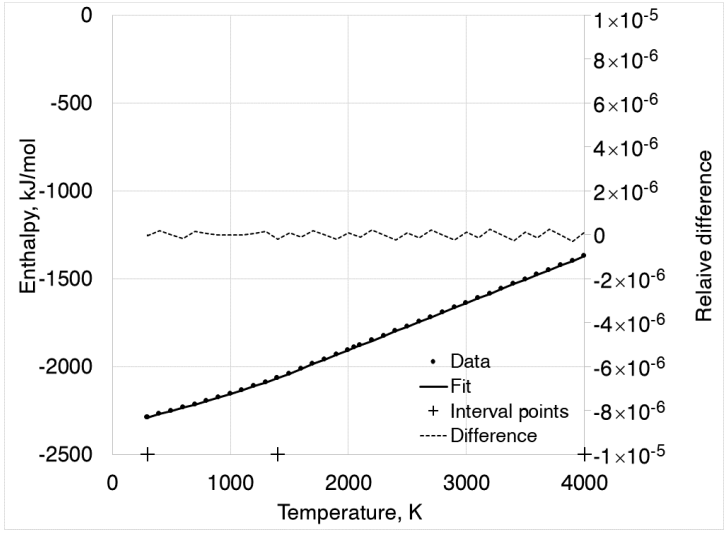


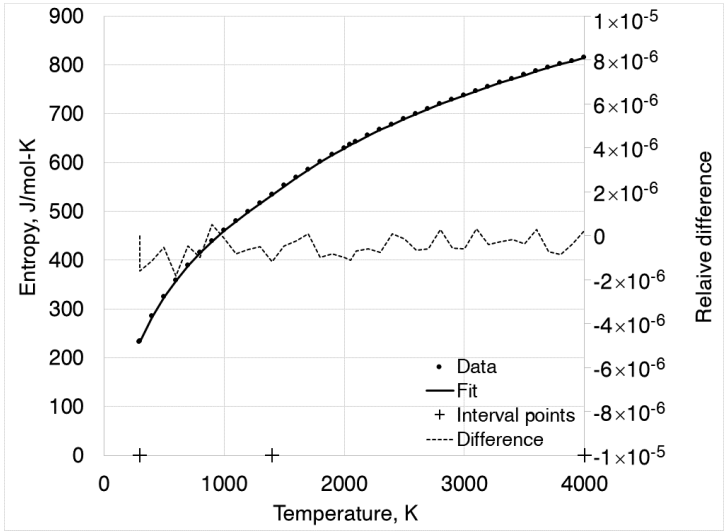


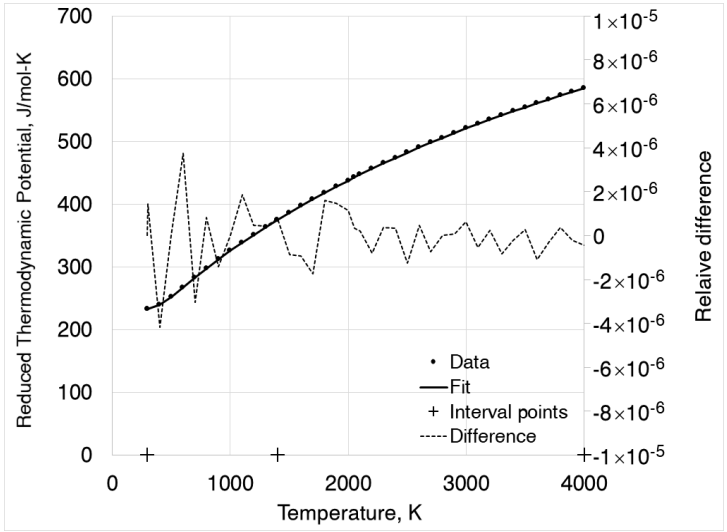
Titanium Oxide (Ti_3O_5) – O5Ti3(l)

MW = 223.64 g/mol – $\Delta_f H^o$(298.15 K) = -2289.059 kJ/mol – S^o(298.15 K) = 232.446 J/mol-K – p^o = 0.1 MPa



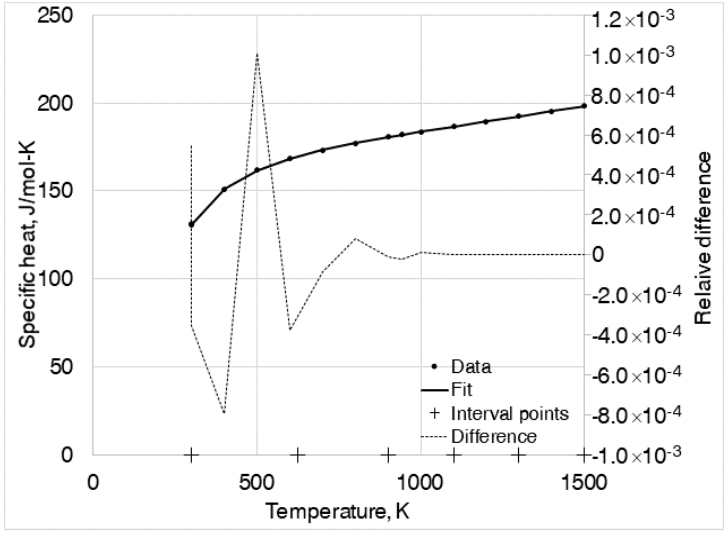


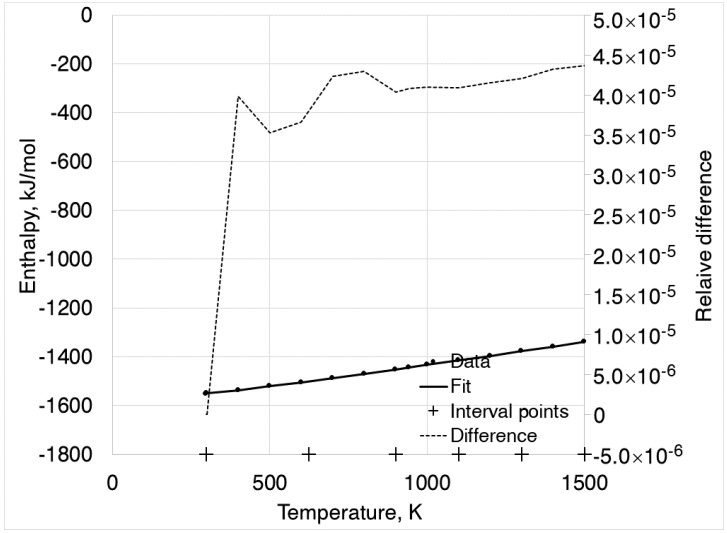


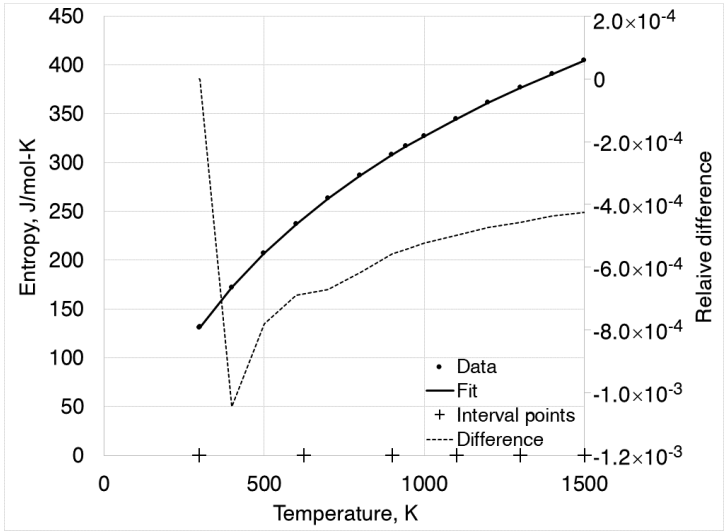


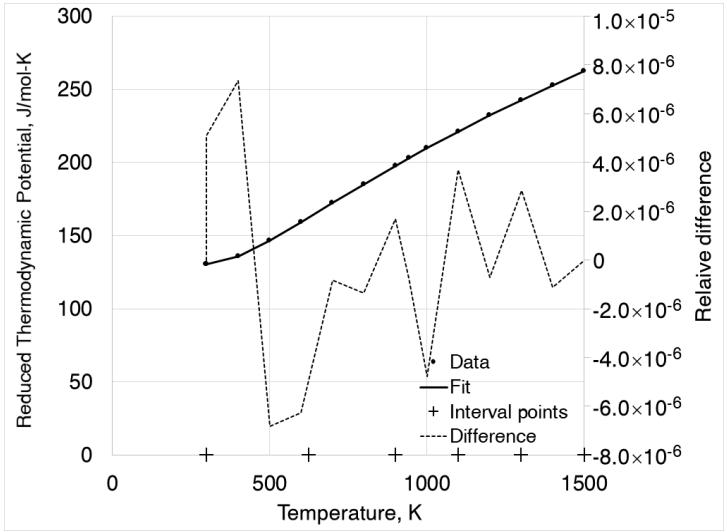
Vanadium Oxide (V_2O_5) – O5V2(cr)

MW = 181.88 g/mol – $\Delta_f H^o$(298.15 K) = -1550.59 kJ/mol – S^o(298.15 K) = 130.549 J/mol-K – p^o = 0.1 MPa



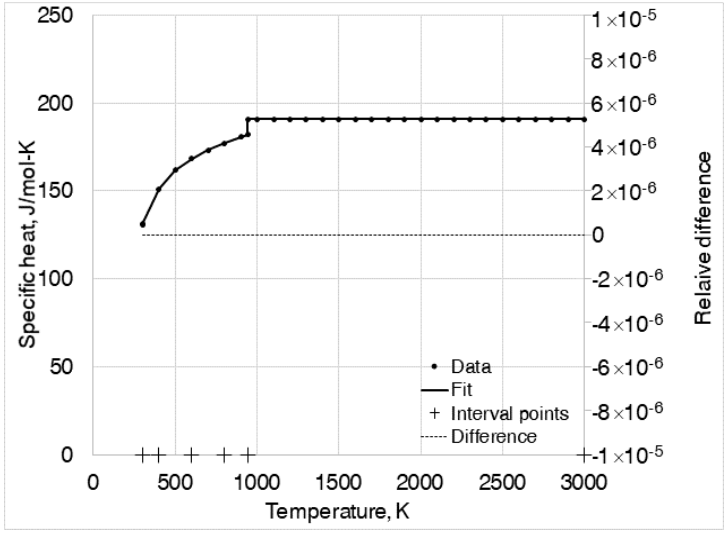


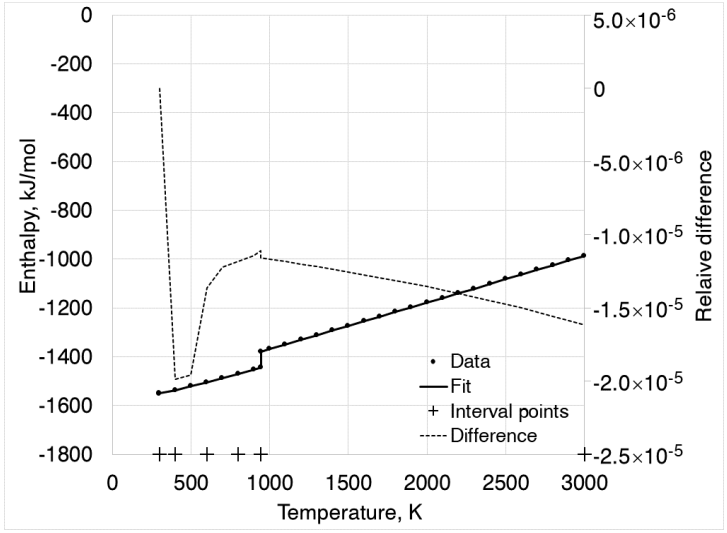


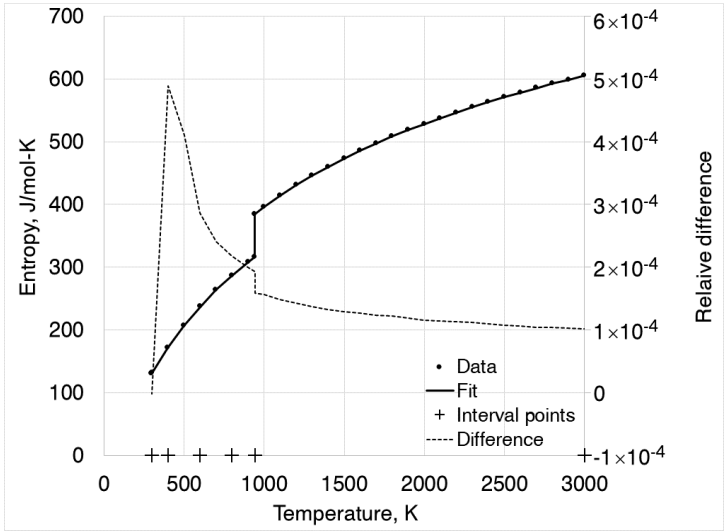


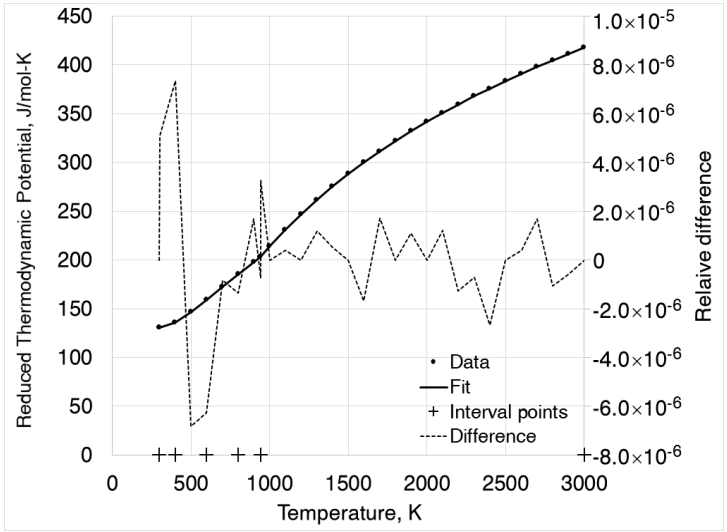
Vanadium Oxide (V_2O_5) – O5V2(cr,l)

MW = 181.88 g/mol – $\Delta_f H^o$(298.15 K) = -1550.59 kJ/mol – S^o(298.15 K) = 130.549 J/mol-K – p^o = 0.1 MPa



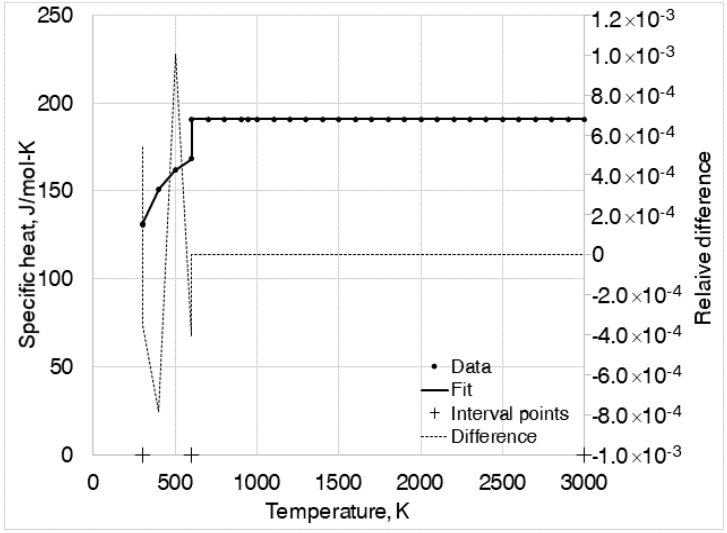


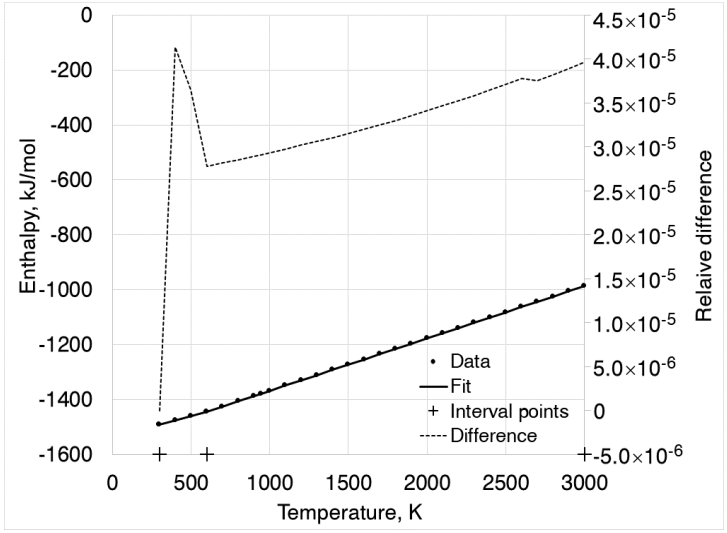


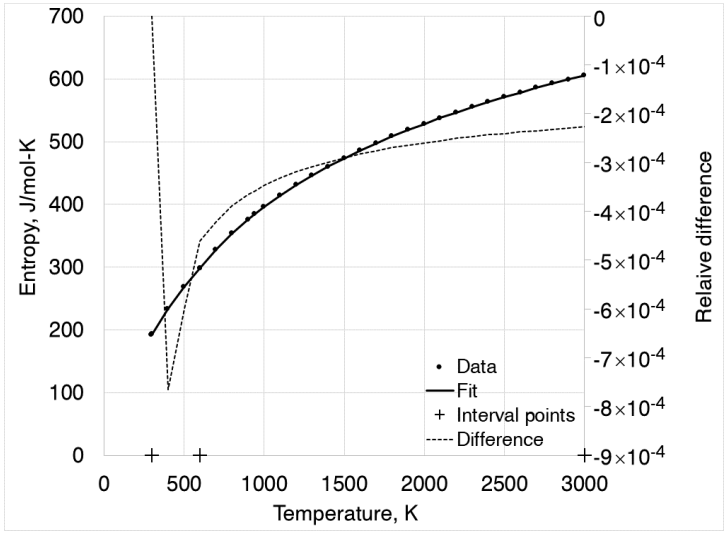


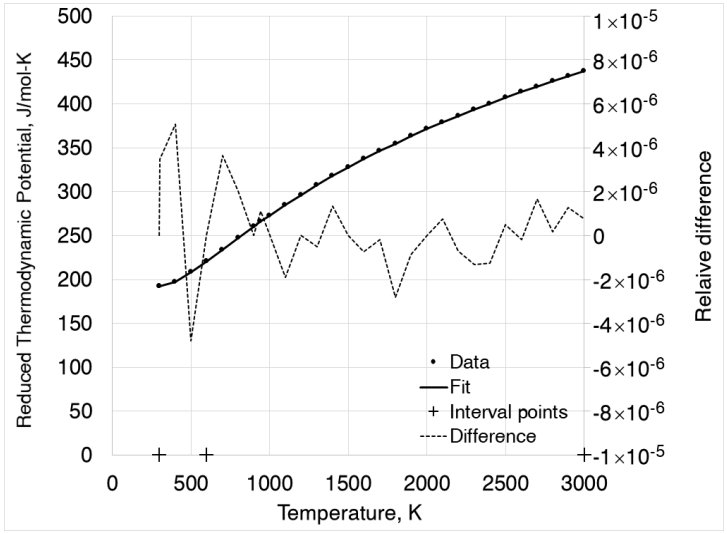
Vanadium Oxide (V_2O_5) – O5V2(l)

MW = 181.88 g/mol – $\Delta_f H^o$(298.15 K) = -1491.209 kJ/mol – S^o(298.15 K) = 191.956 J/mol-K – p^o = 0.1 MPa



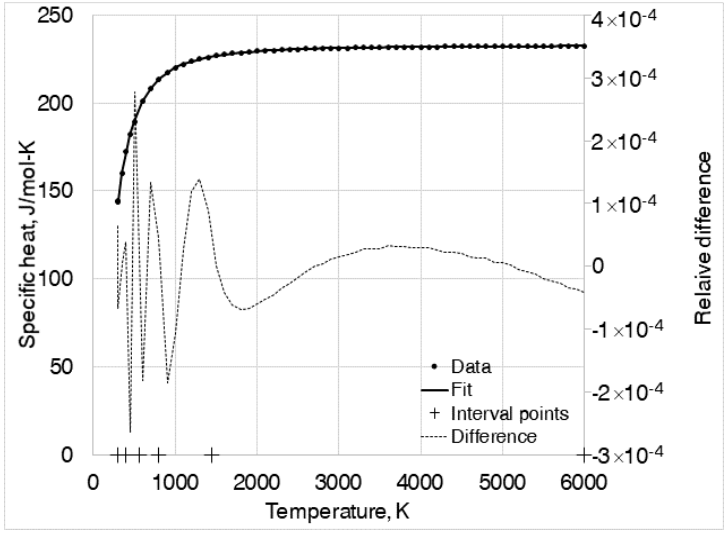


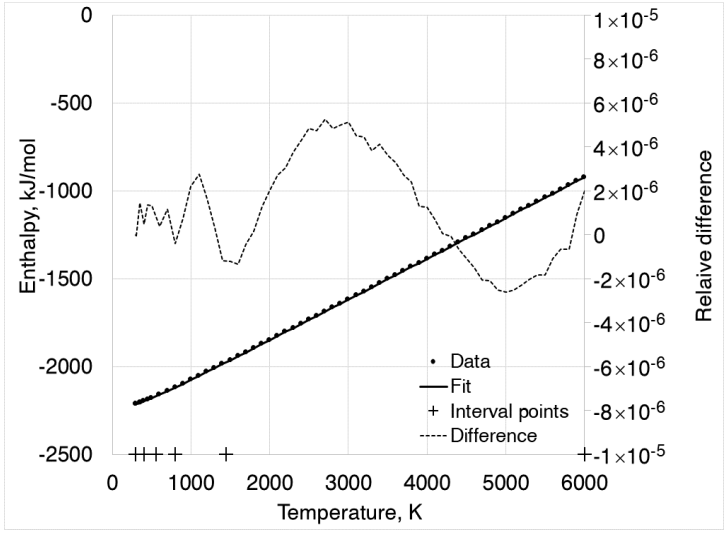


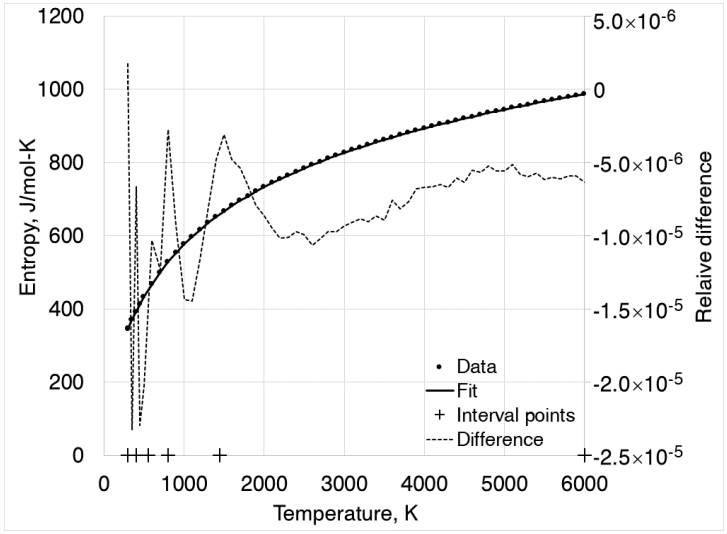


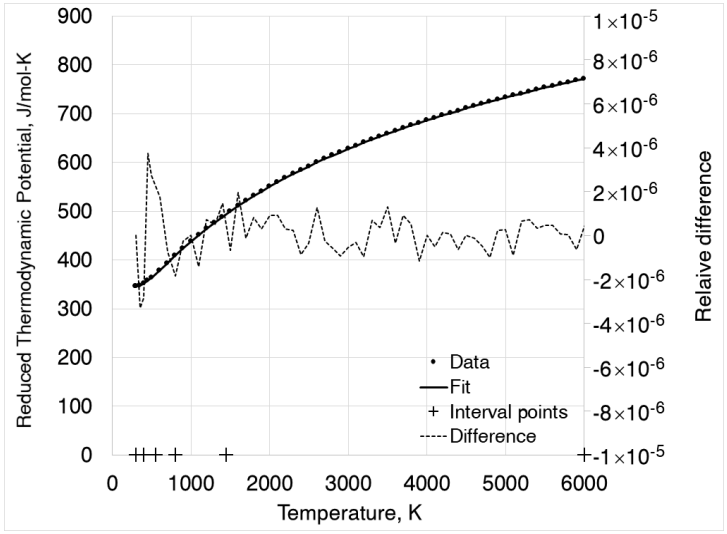
Phosphorus Oxide ($(P_2O_3)_2$) – O6P4(g)

MW = 219.8914 g/mol – $\Delta_f H^o$(298.15 K) = -2214.307 kJ/mol – S^o(298.15 K) = 345.641 J/mol-K – p^o = 0.1 MPa



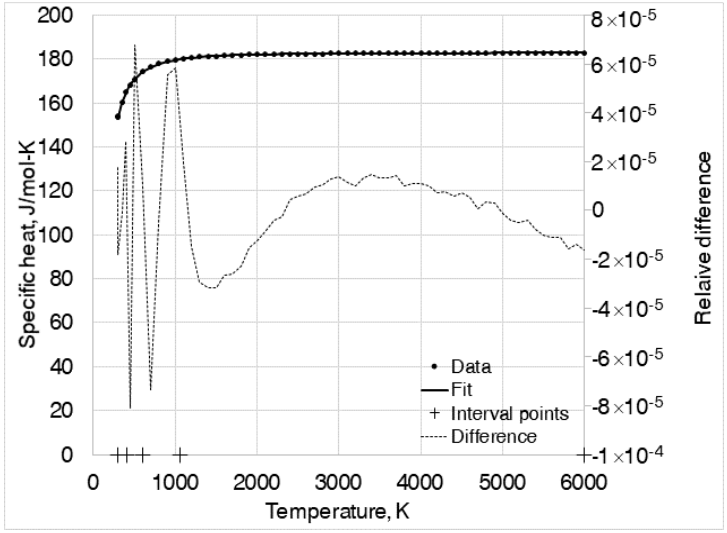


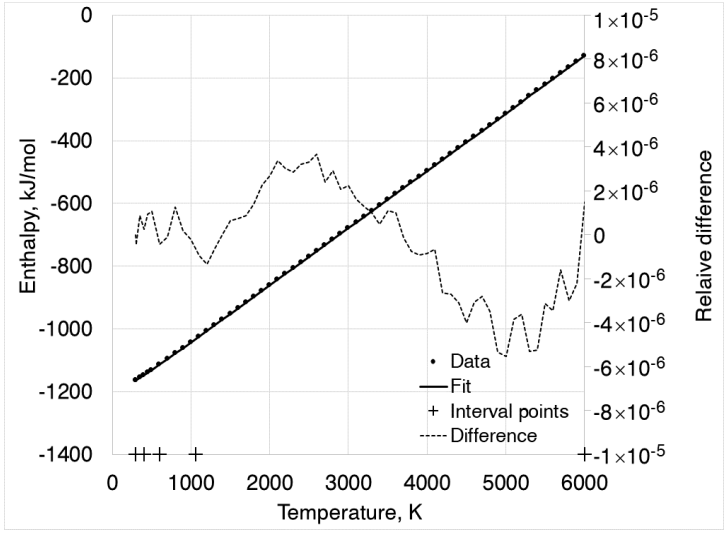


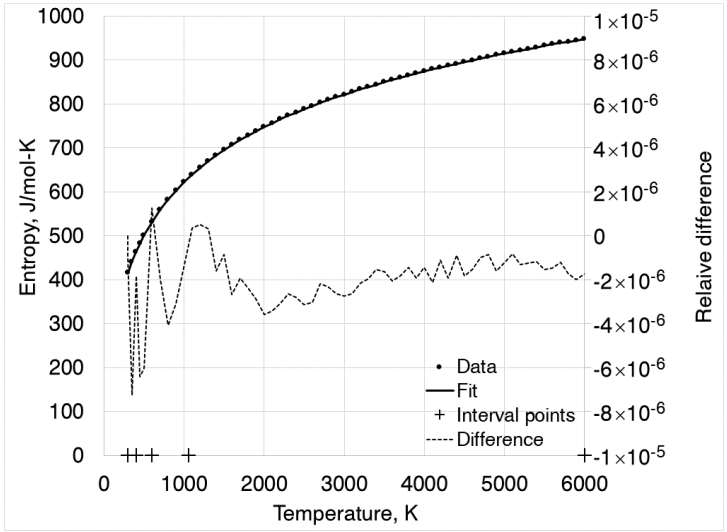


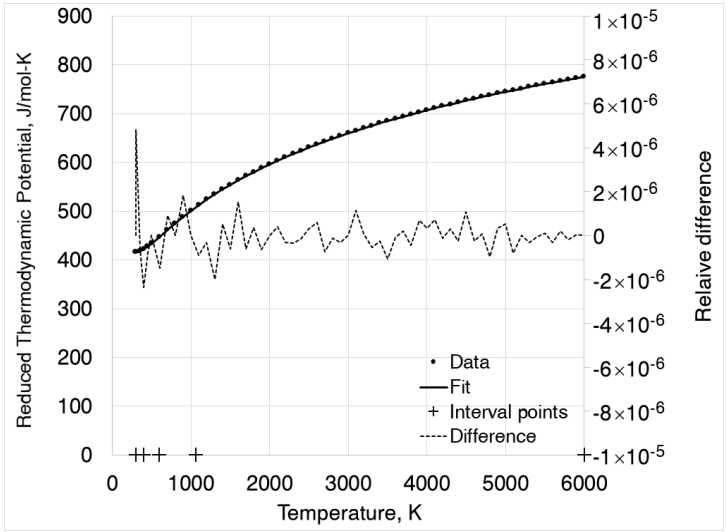
Tungsten Oxide ($(WO_3)_2$) – O6W2(g)

MW = 463.7 g/mol – $\Delta_f H^o$(298.15 K) = -1163.989 kJ/mol – S^o(298.15 K) = 415.601 J/mol-K – p^o = 0.1 MPa



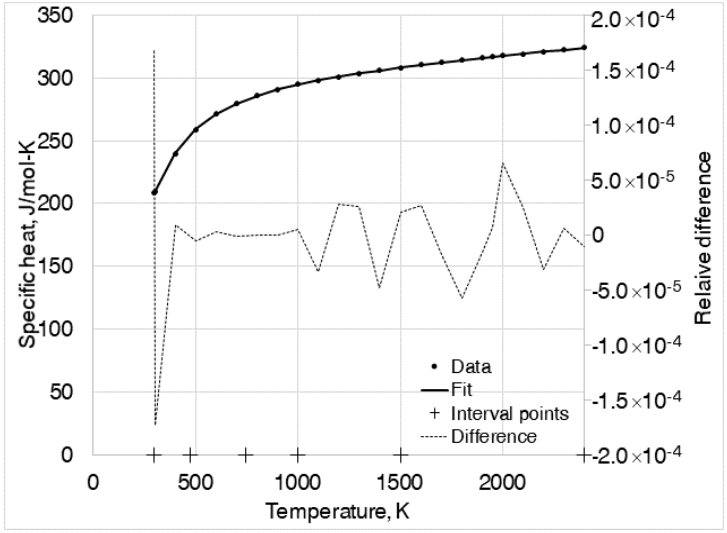


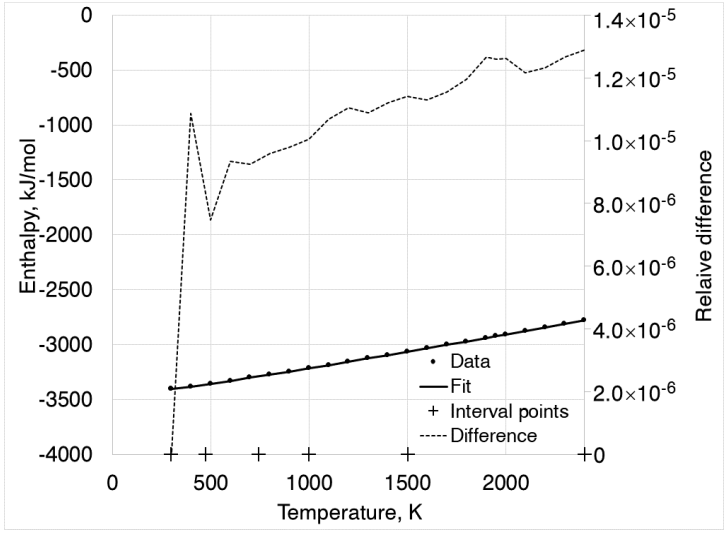


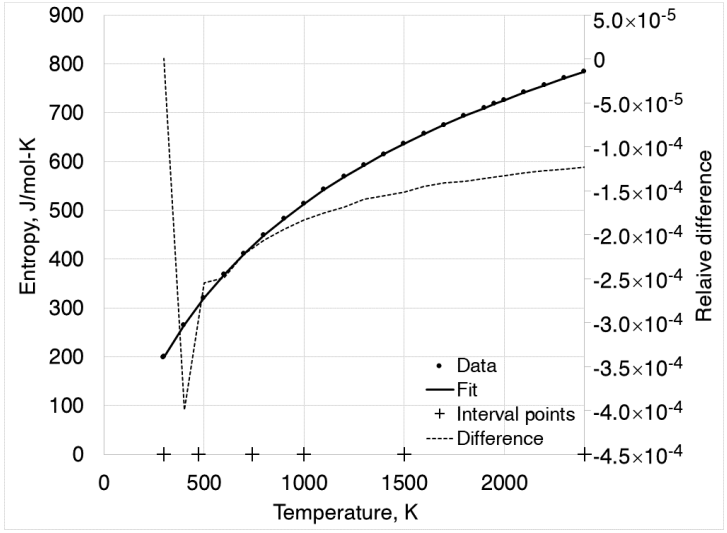


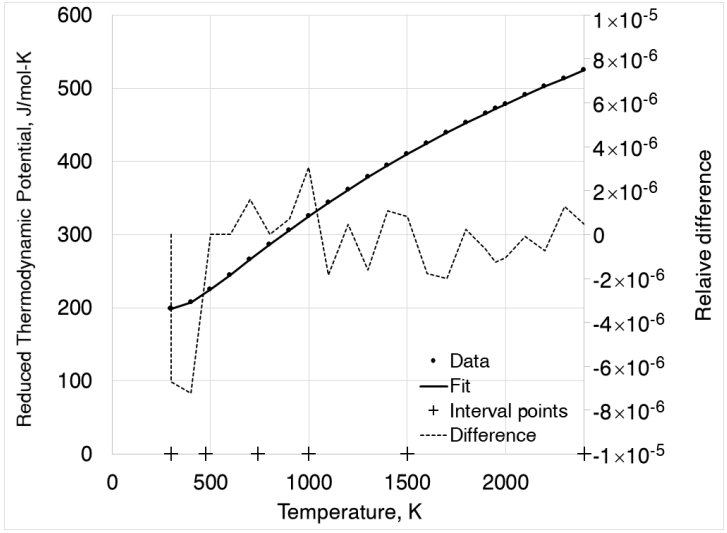
Titanium Oxide (Ti_4O_7) – O7Ti4(cr)

MW = 303.52 g/mol – $\Delta_f H^o$(298.15 K) = -3404.521 kJ/mol – S^o(298.15 K) = 198.74 J/mol-K – p^o = 0.1 MPa



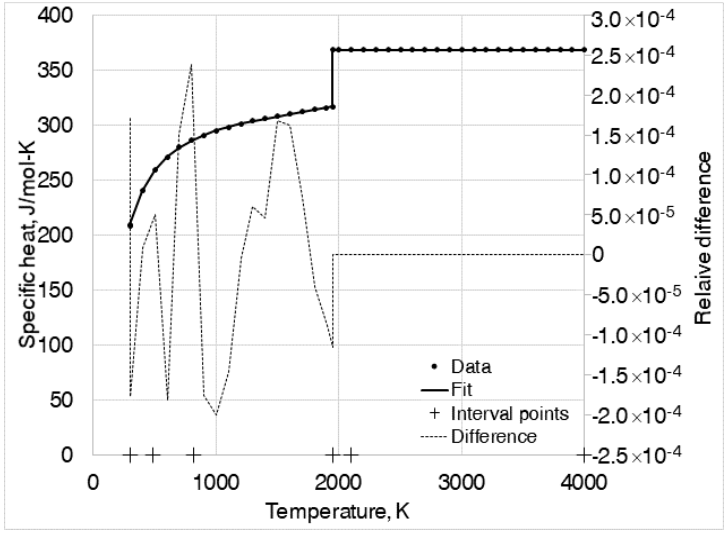


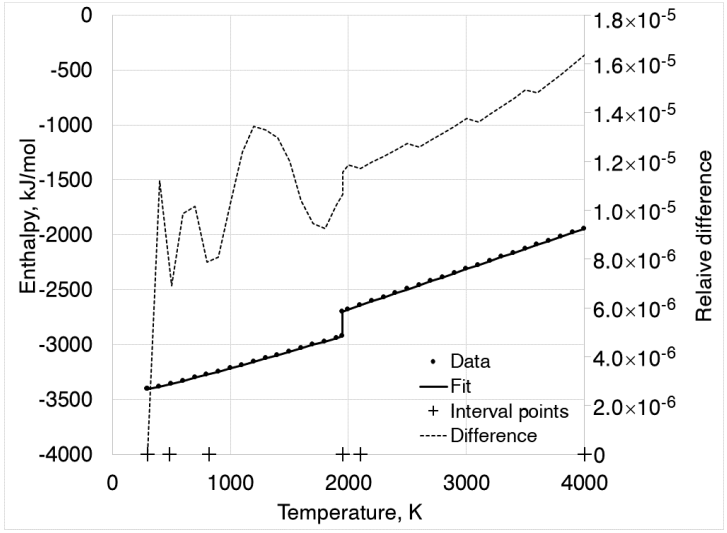


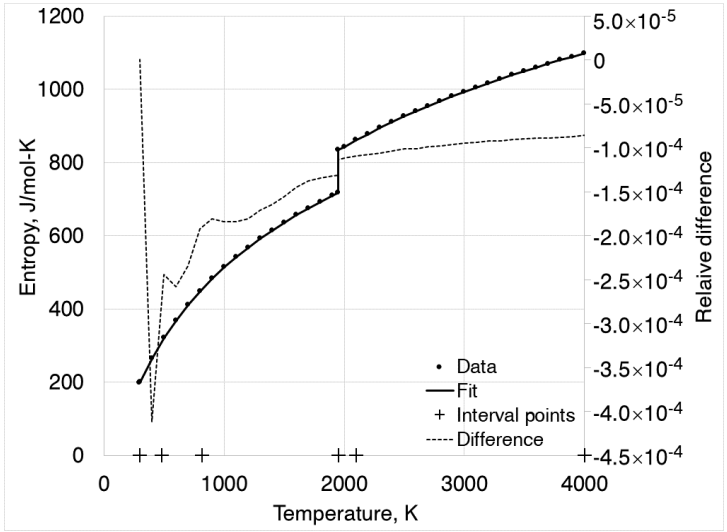


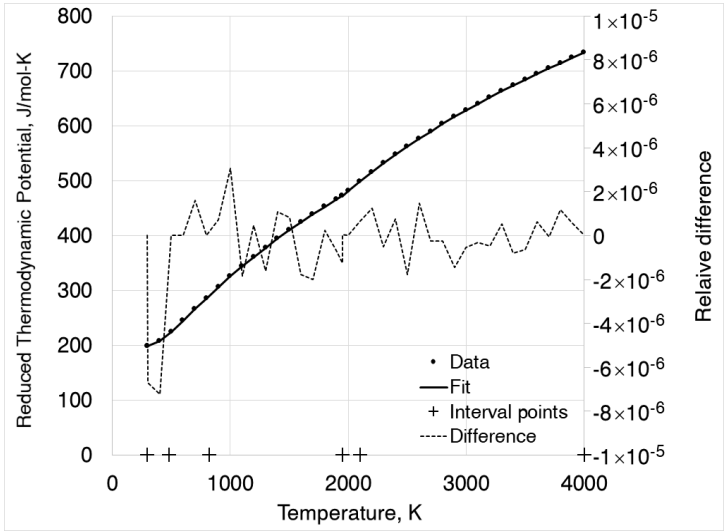
Titanium Oxide (Ti_4O_7) – O7Ti4(cr,l)

MW = 303.52 g/mol – $\Delta_f H^o$(298.15 K) = -3404.521 kJ/mol – S^o(298.15 K) = 198.74 J/mol-K – p^o = 0.1 MPa



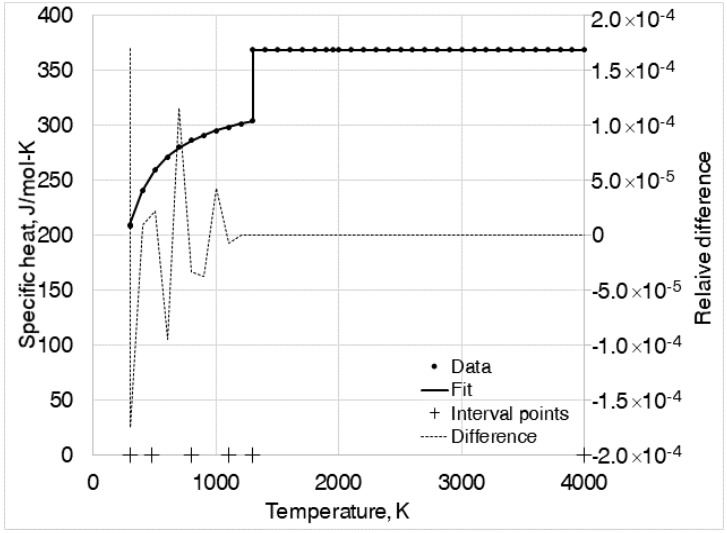


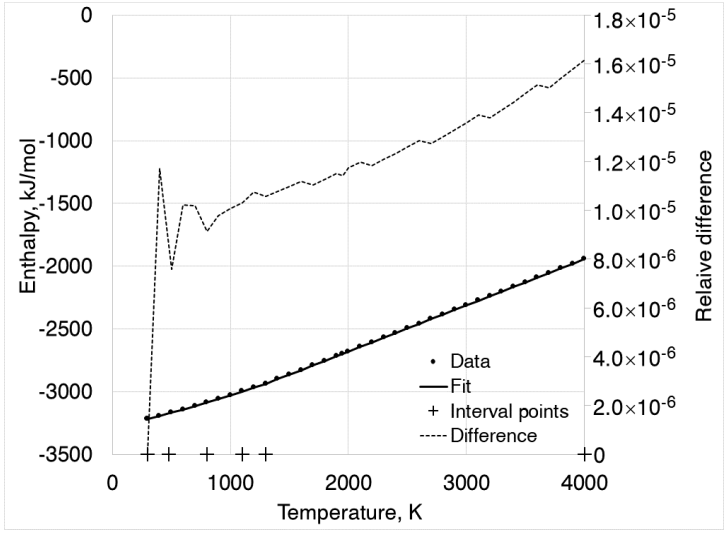


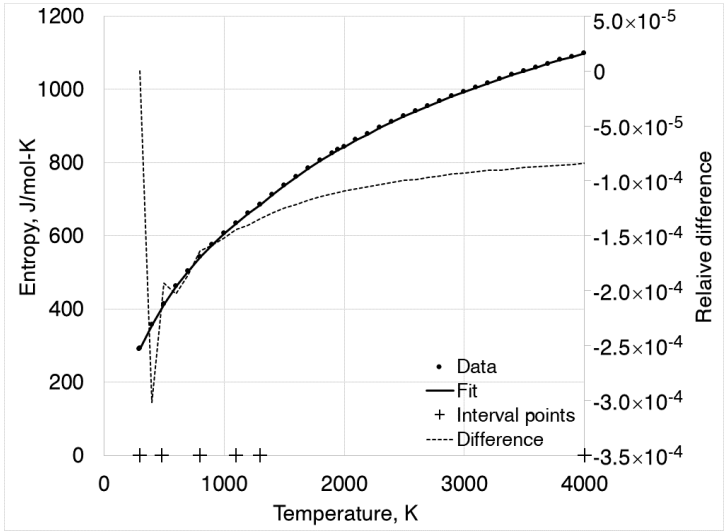


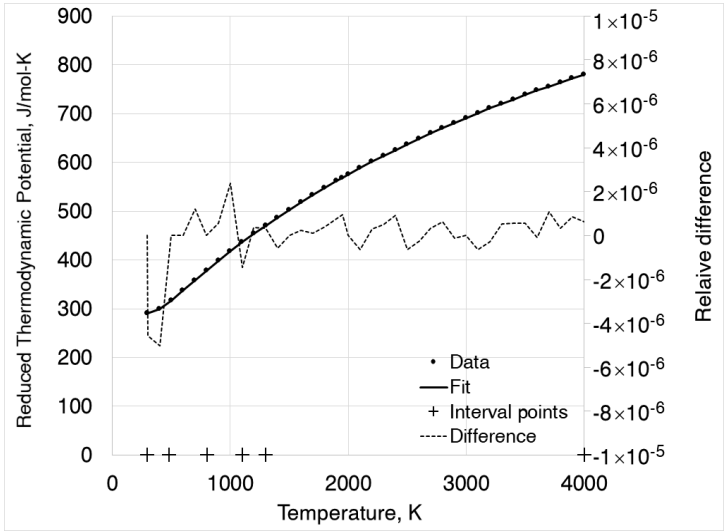
Titanium Oxide (Ti_4O_7) – O7Ti4(l)

MW = 303.52 g/mol – $\Delta_f H^o$(298.15 K) = -3216.035 kJ/mol – S^o(298.15 K) = 291.066 J/mol-K – p^o = 0.1 MPa



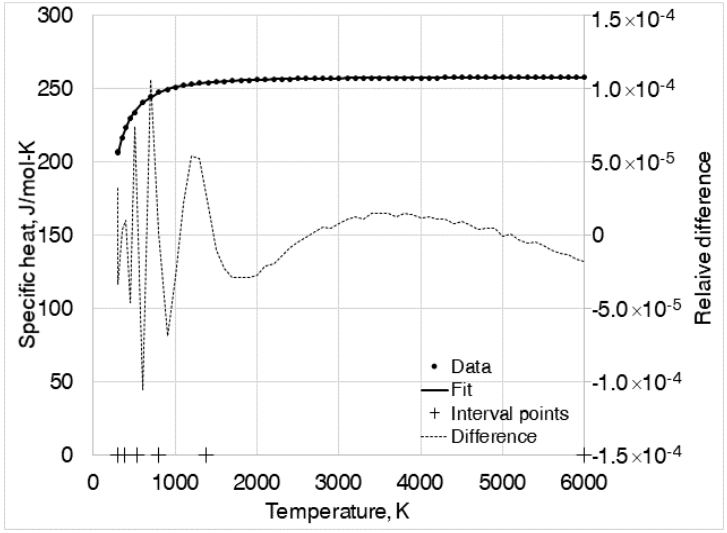


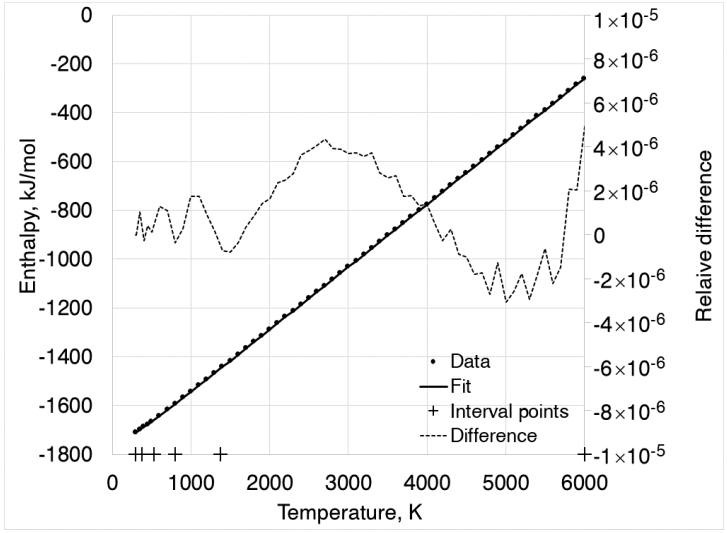


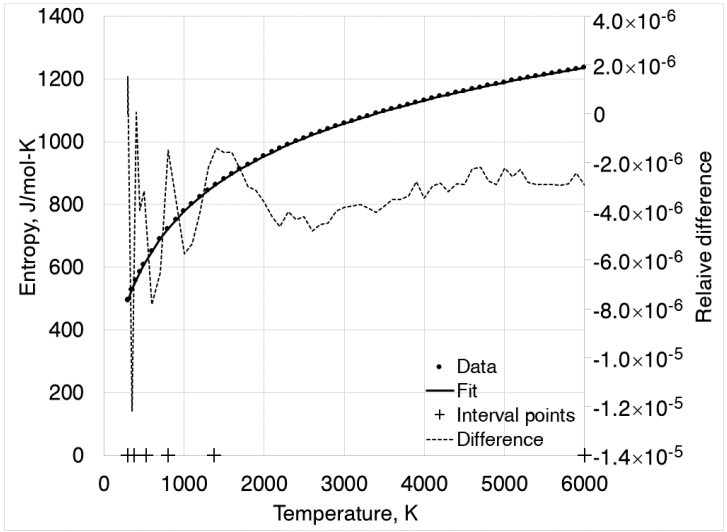


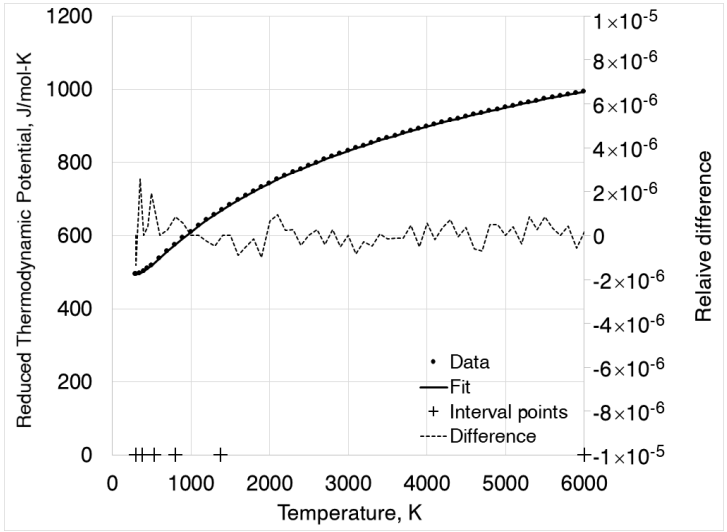
Tungsten Oxide (W_3O_8) – O8W3(g)

MW = 679.55 g/mol – $\Delta_f H^o$(298.15 K) = -1710.001 kJ/mol – S^o(298.15 K) = 494.013 J/mol-K – p^o = 0.1 MPa



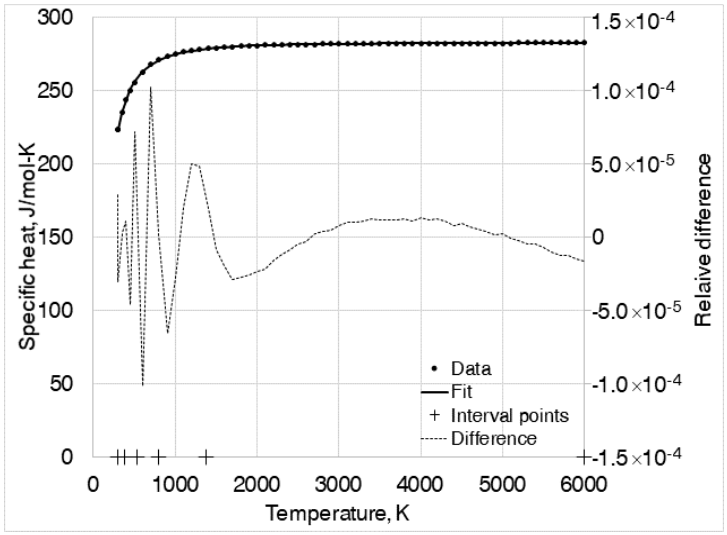


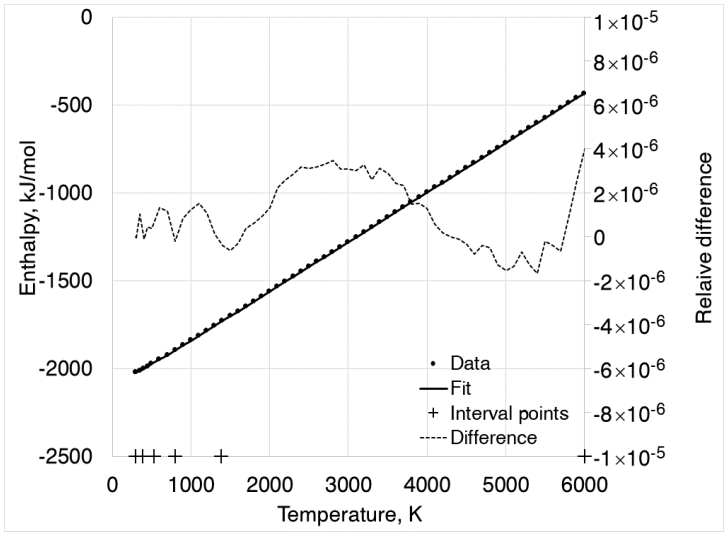


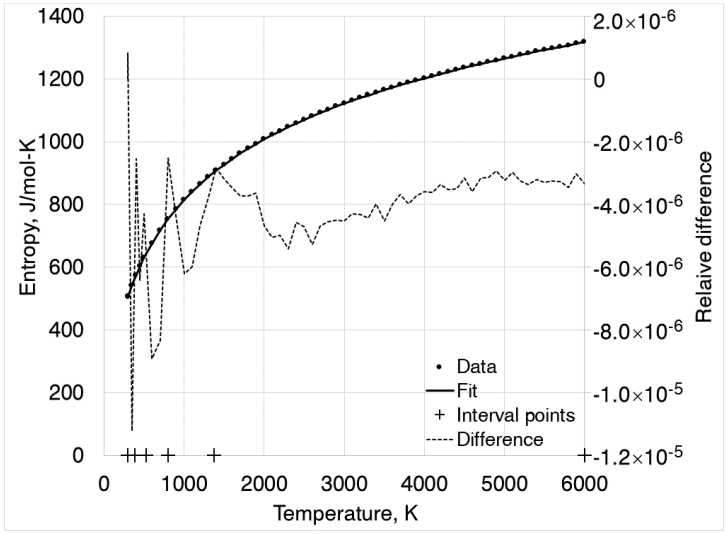


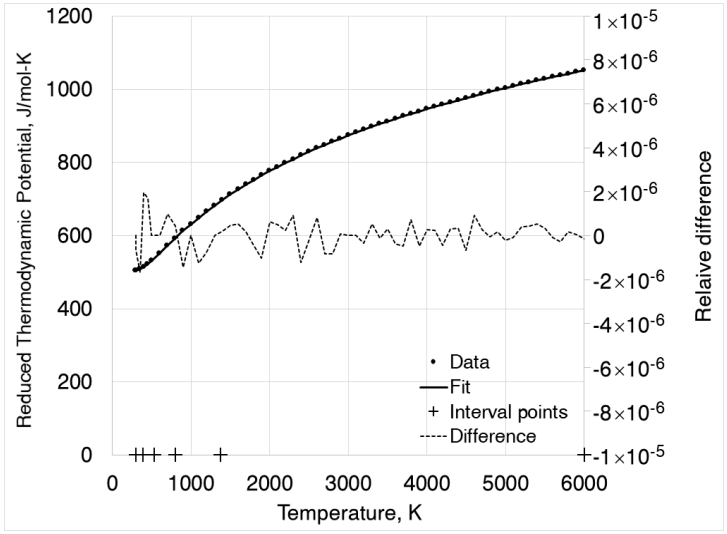
Tungsten Oxide ($(WO_3)_3$) – O9W3(g)

MW = 695.54 g/mol – $\Delta_f H^o$(298.15 K) = -2023.382 kJ/mol – S^o(298.15 K) = 504.674 J/mol-K – p^o = 0.1 MPa



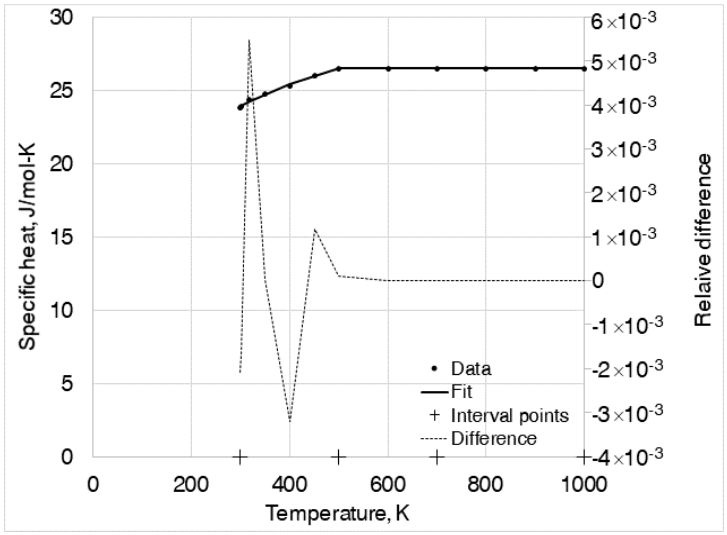


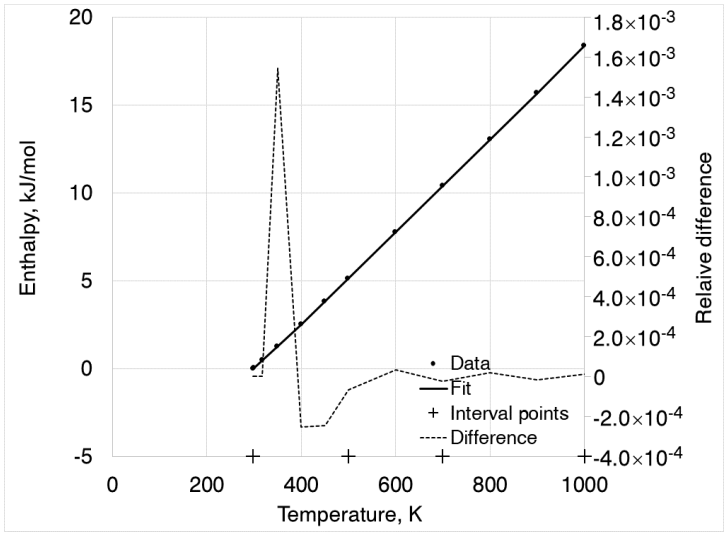


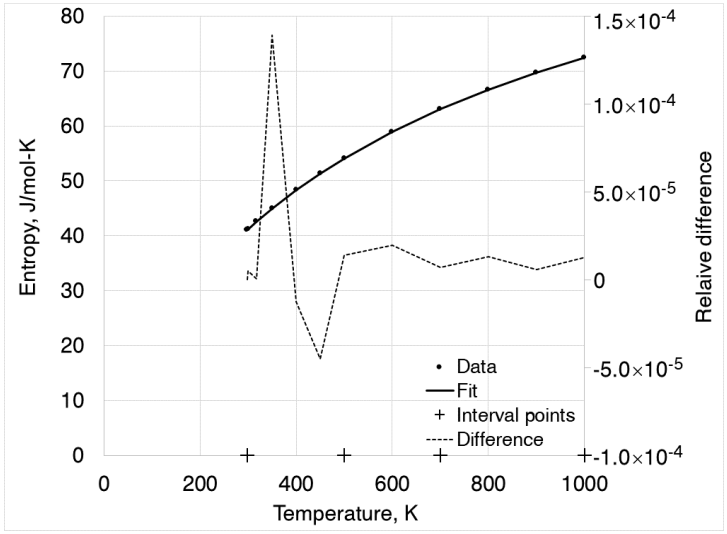


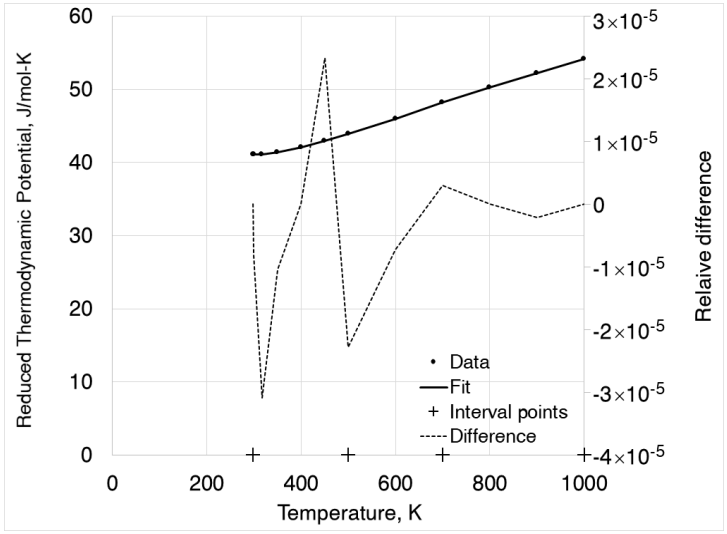
Phosphorus, White (P) – P1(cr)_01

MW = 30.97376 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 41.077 J/mol-K – p^o = 0.1 MPa



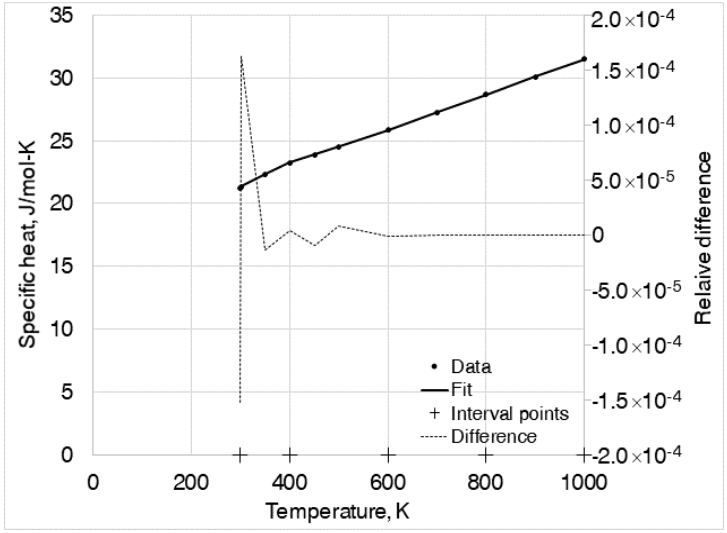


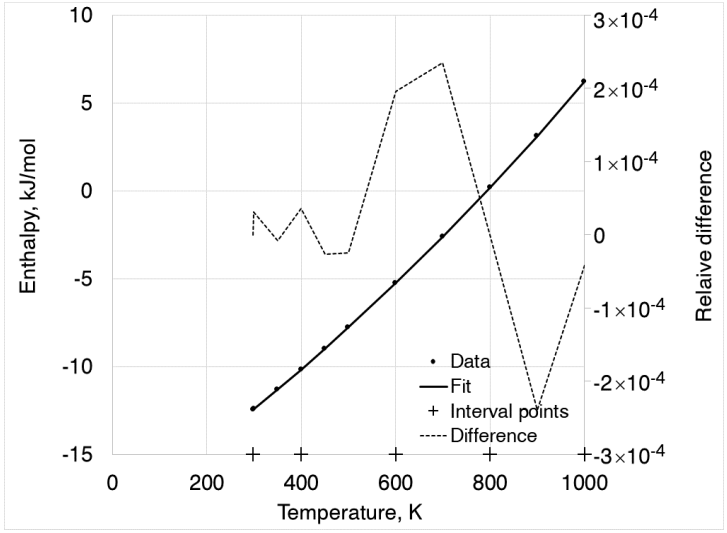


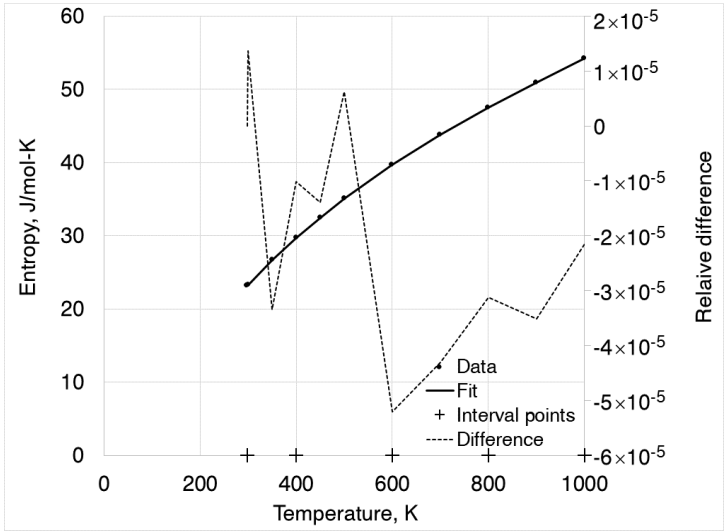


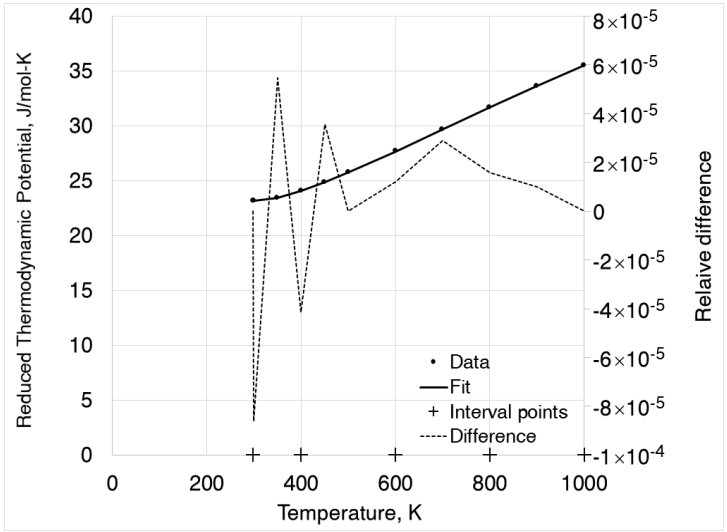
Phosphorus, Red, IV (P) – P1(cr)_02

MW = 30.97376 g/mol – $\Delta_f H^o$(298.15 K) = -12.439 kJ/mol – S^o(298.15 K) = 23.197 J/mol-K – p^o = 0.1 MPa



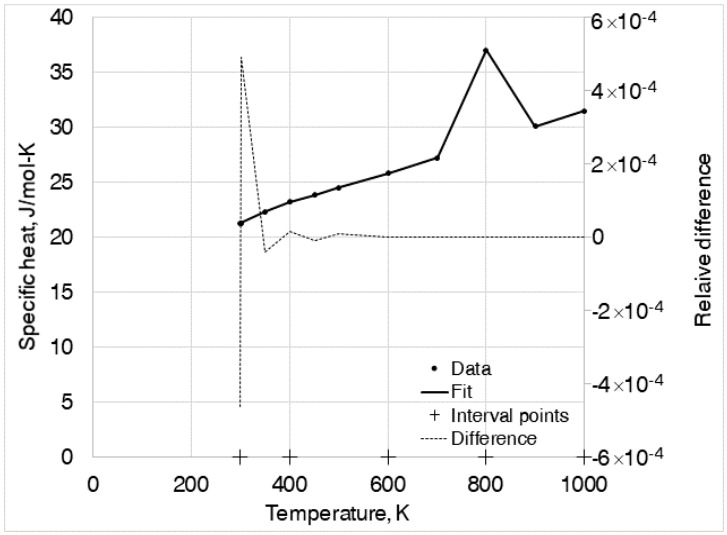


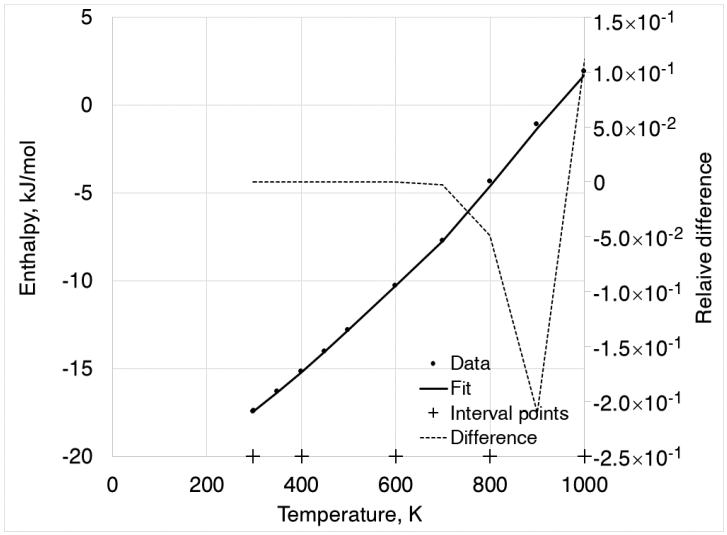


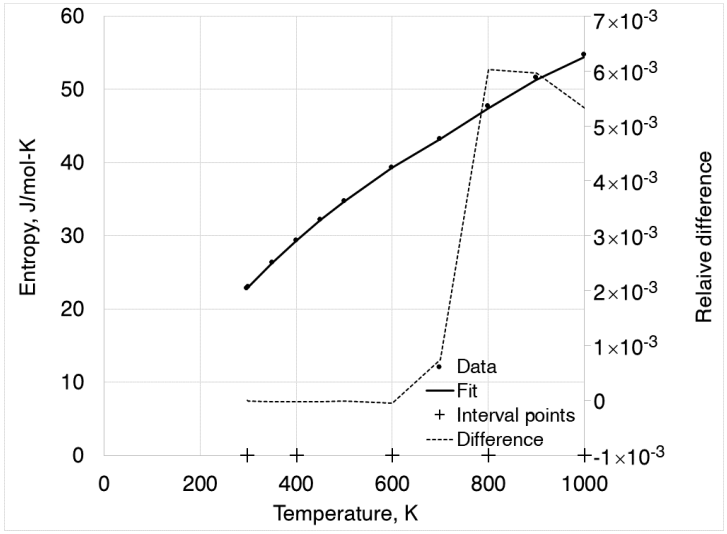


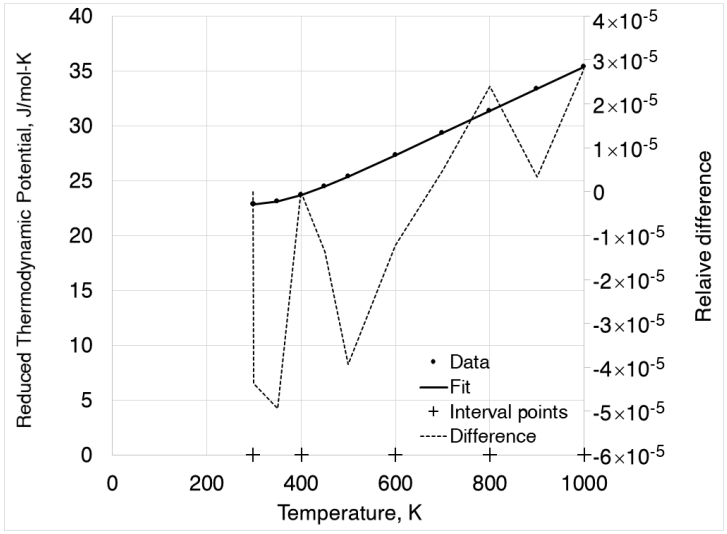
Phosphorus, Red, V (P) – P1(cr)_03

MW = 30.97376 g/mol – $\Delta_f H^o$(298.15 K) = -17.46 kJ/mol – S^o(298.15 K) = 22.853 J/mol-K – p^o = 0.1 MPa



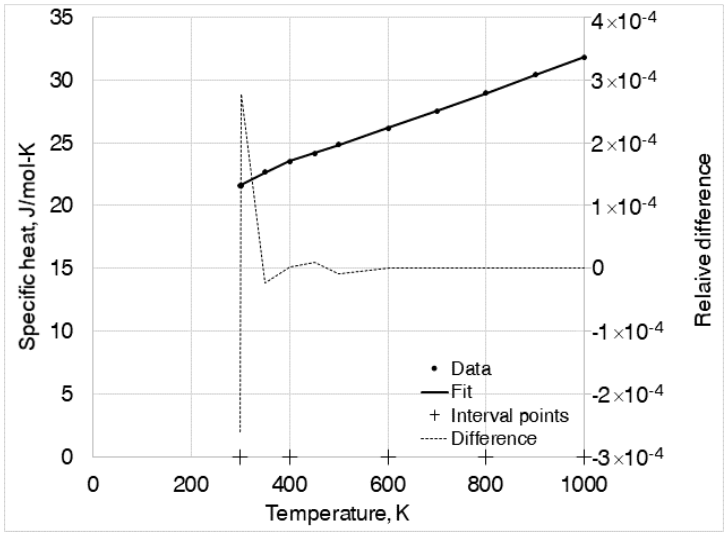


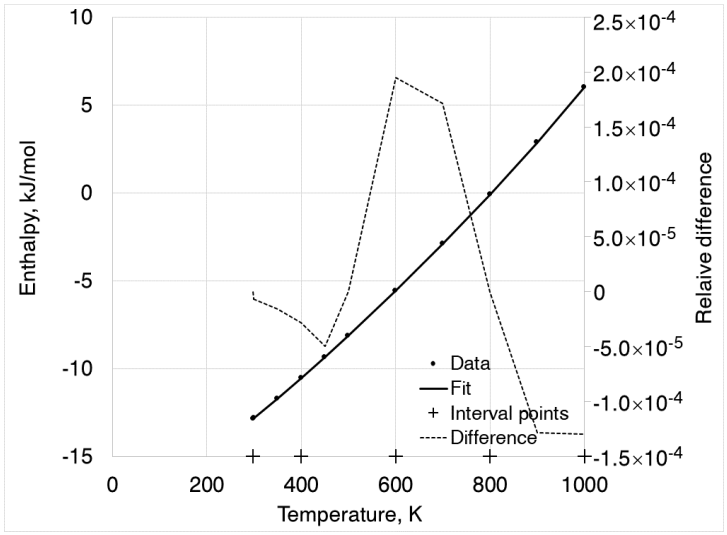


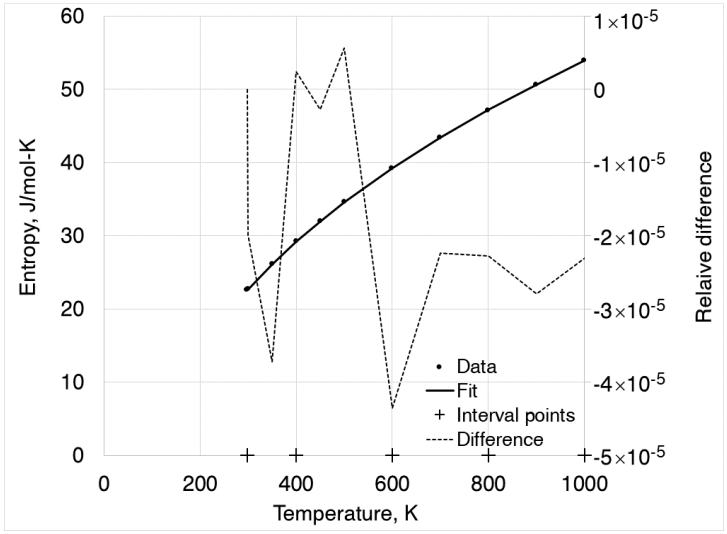


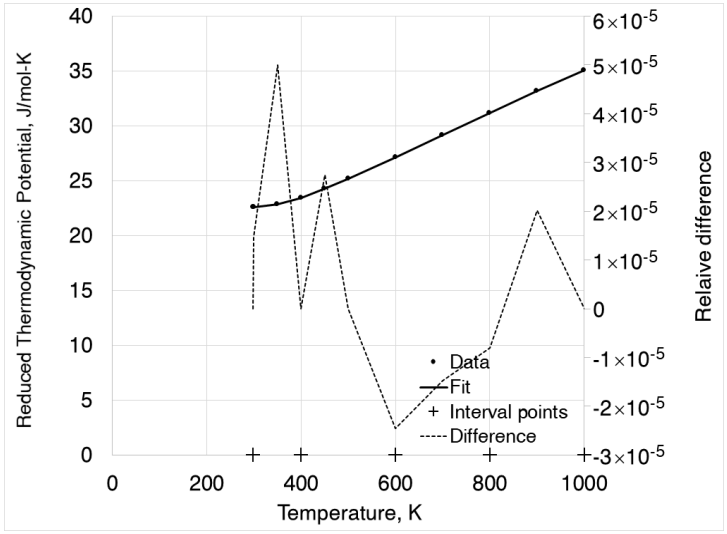
Phosphorus, Black (P) – P1(cr)_04

MW = 30.97376 g/mol – $\Delta_f H^o$(298.15 K) = -12.851 kJ/mol – S^o(298.15 K) = 22.586 J/mol-K – p^o = 0.1 MPa



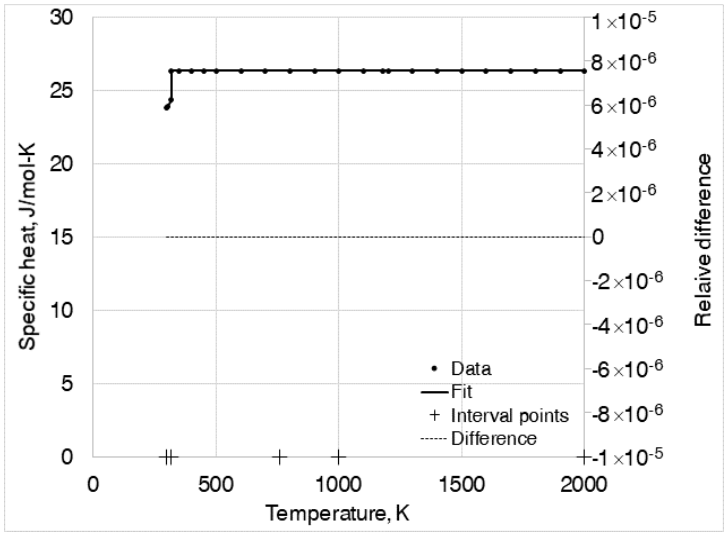


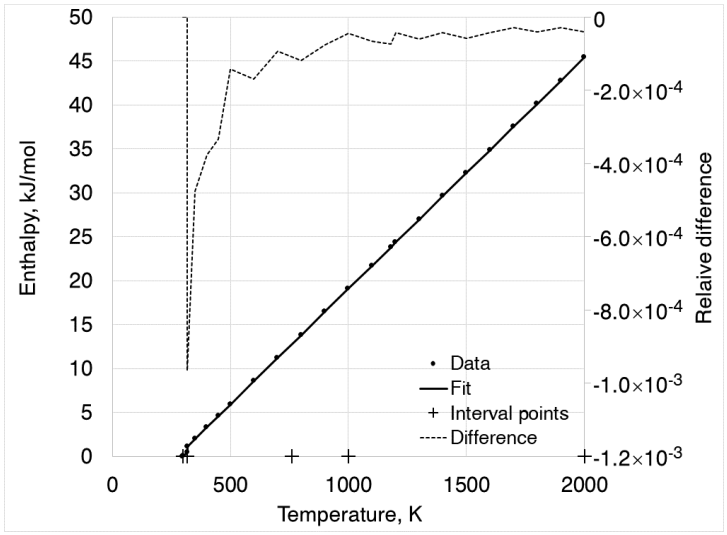


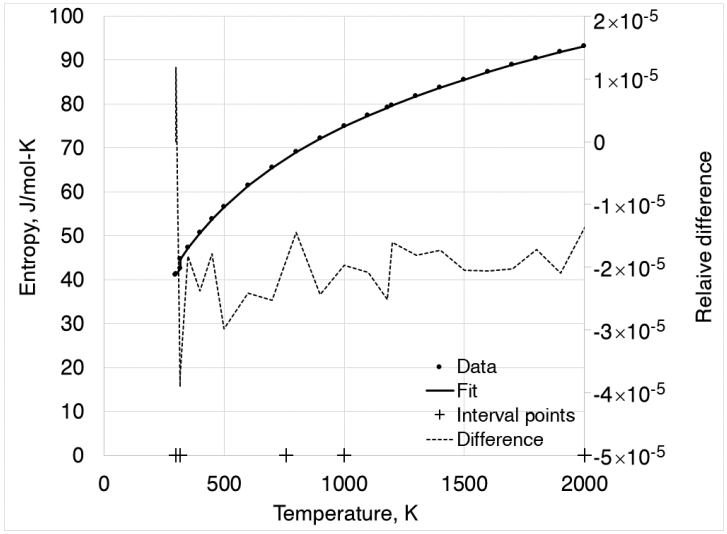


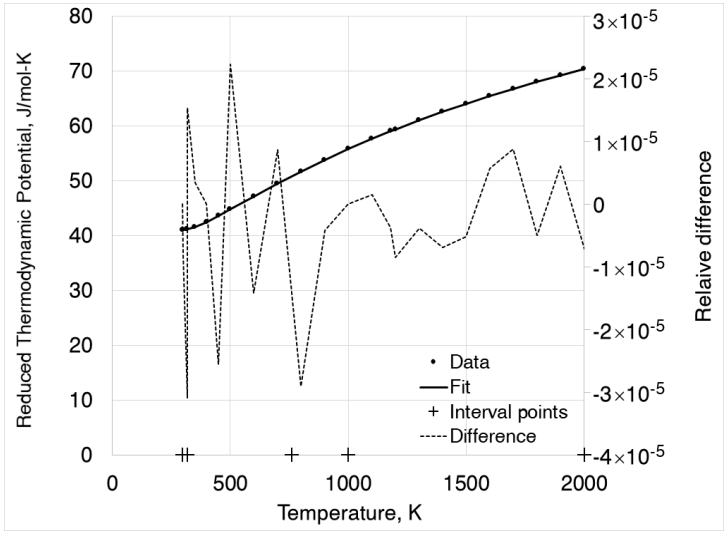
Phosphorus (P) – P1(cr,l)

MW = 30.97376 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 41.077 J/mol-K – p^o = 0.1 MPa



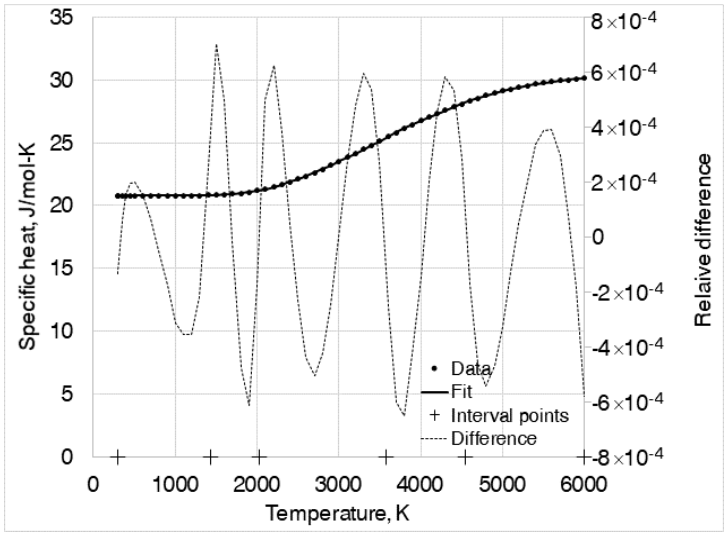


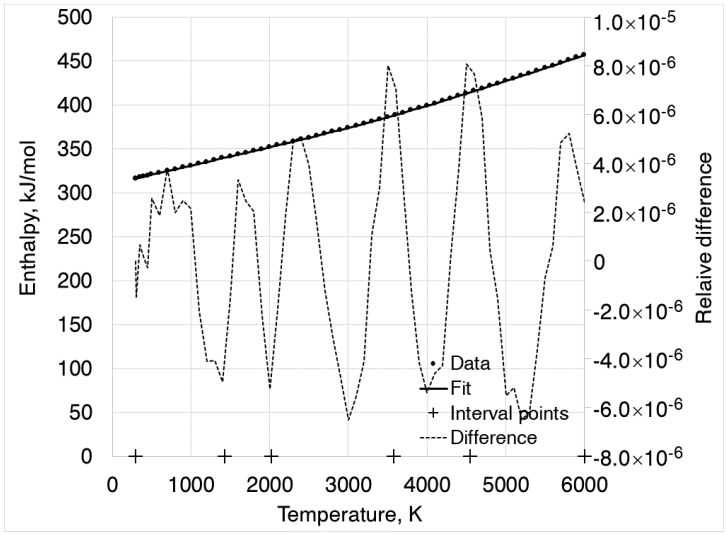


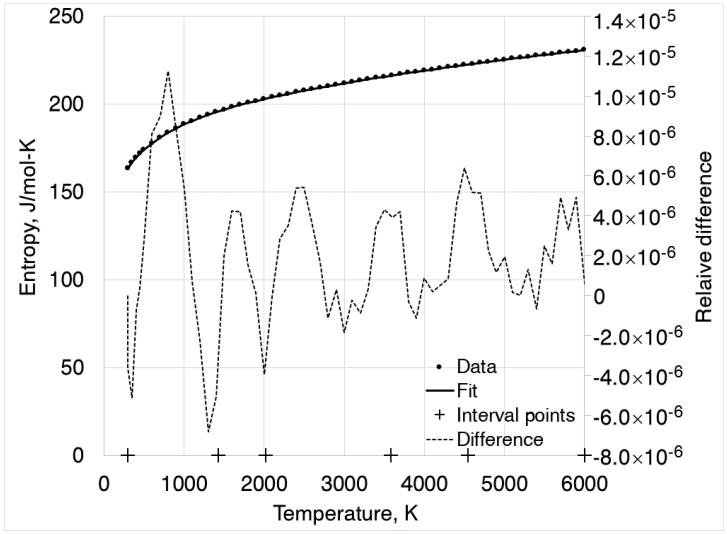


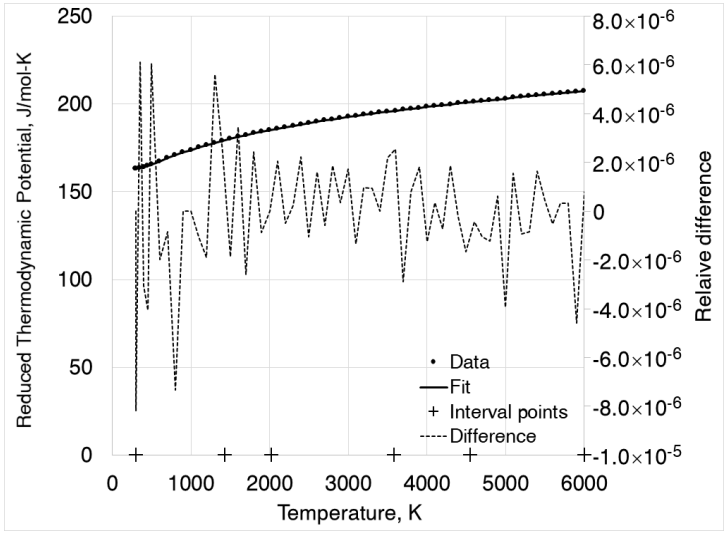
Phosphorus (P) – P1(g)

MW = 30.97376 g/mol – $\Delta_f H^o$(298.15 K) = 316.39 kJ/mol – S^o(298.15 K) = 163.198 J/mol-K – p^o = 0.1 MPa



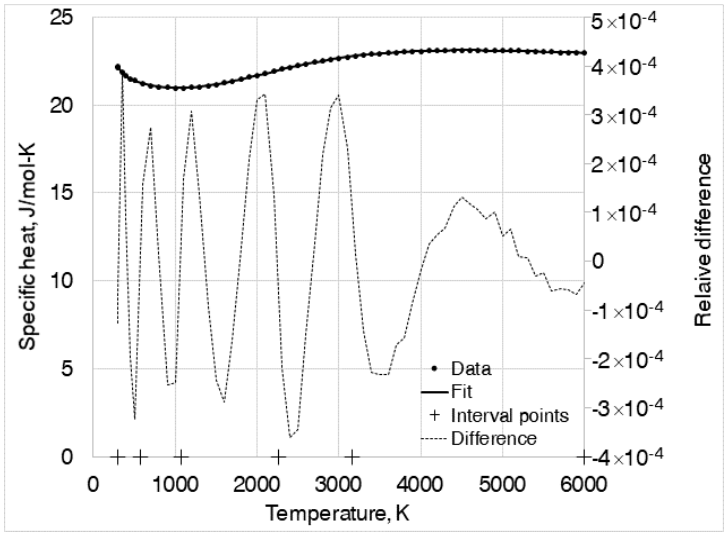


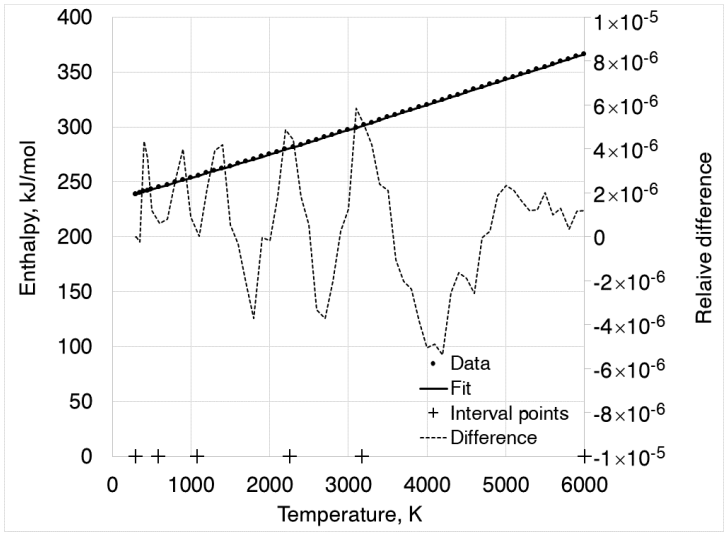


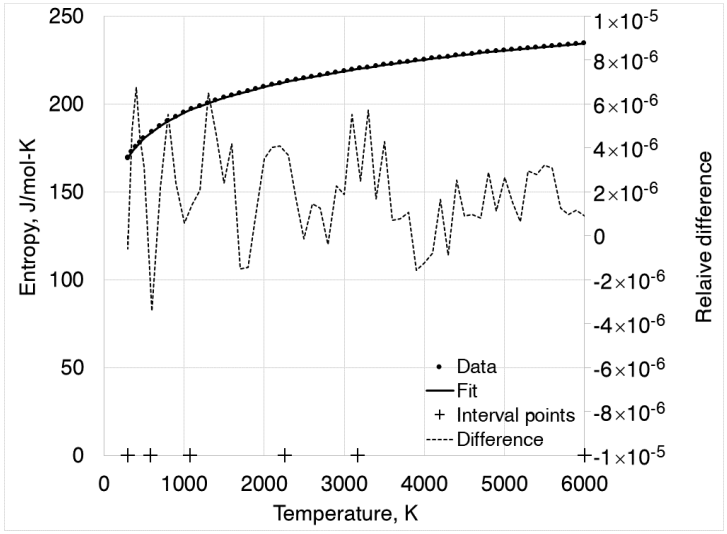


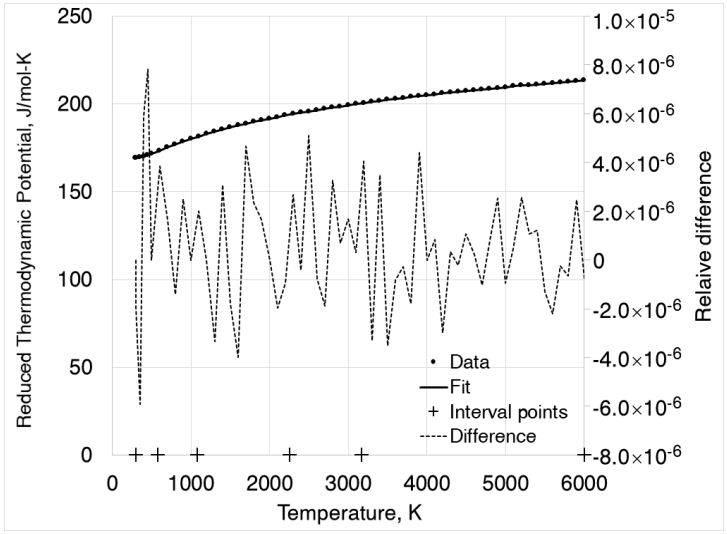
Phosphorus, Ion (P^-) – P1-(g)

MW = 30.97431 g/mol – $\Delta_f H^o$(298.15 K) = 238.717 kJ/mol – S^o(298.15 K) = 169.124 J/mol-K – p^o = 0.1 MPa



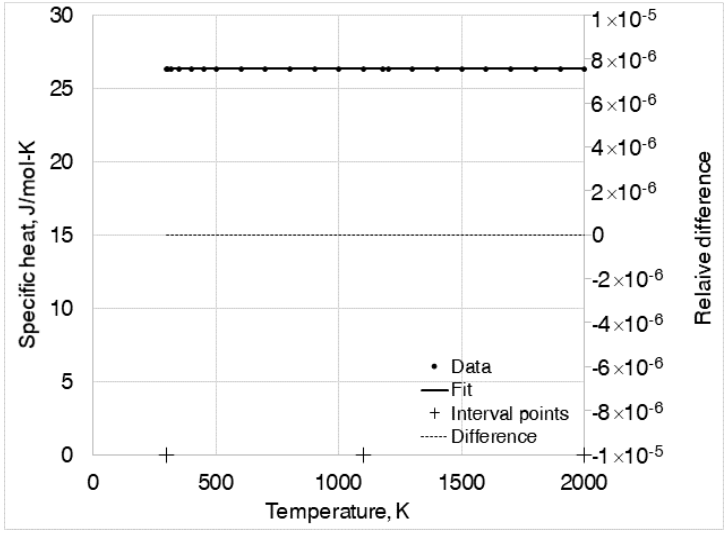


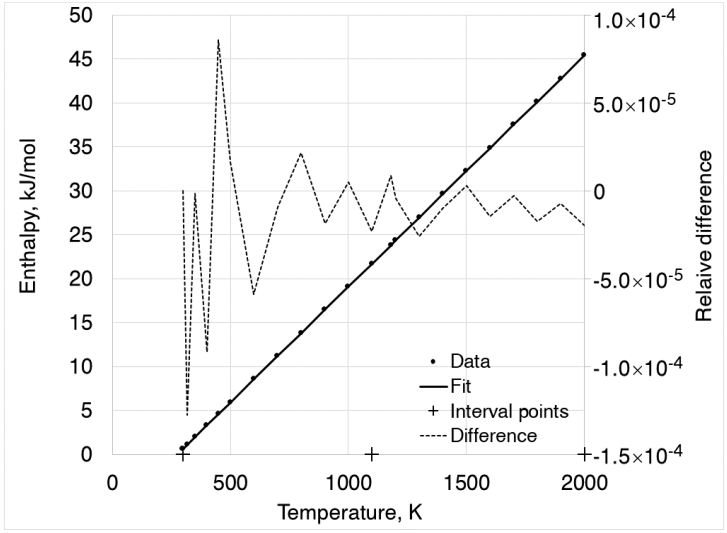


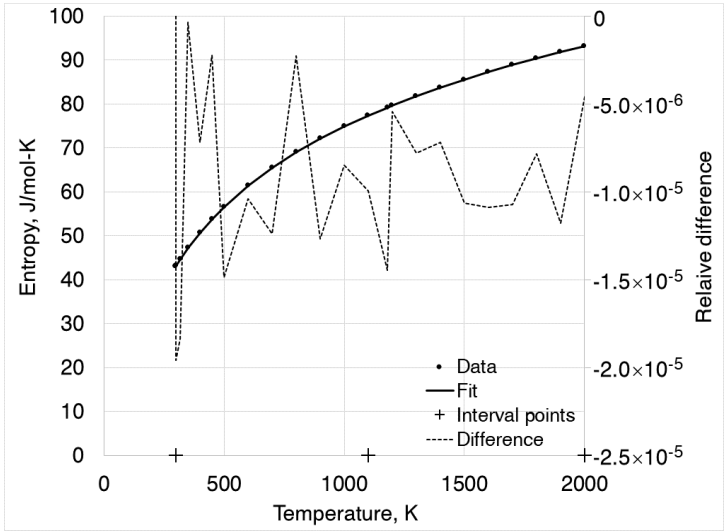


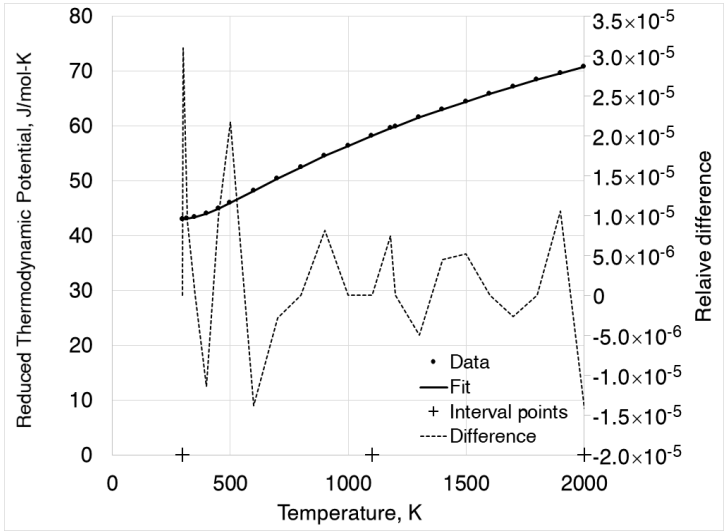
Phosphorus (P) – P1(l)

MW = 30.97376 g/mol – $\Delta_f H^o$(298.15 K) = .615 kJ/mol – S^o(298.15 K) = 43.011 J/mol-K – p^o = 0.1 MPa



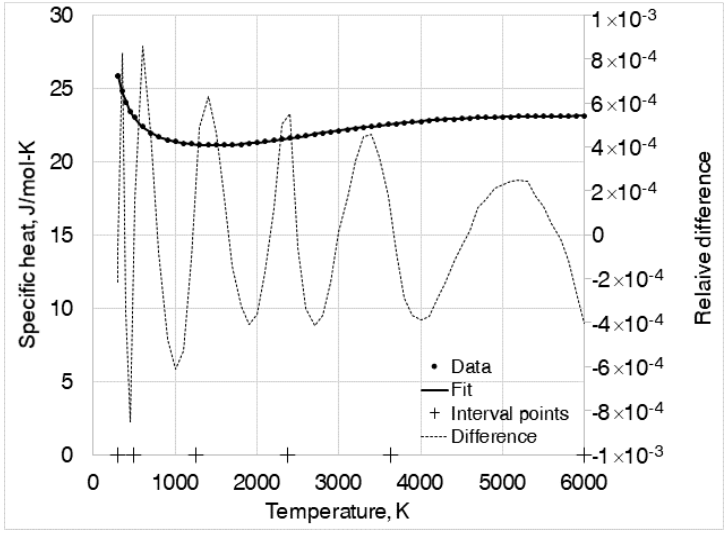


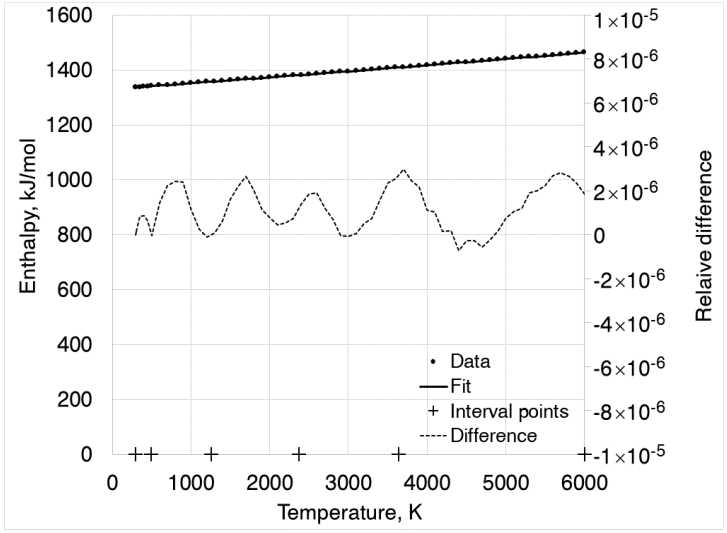


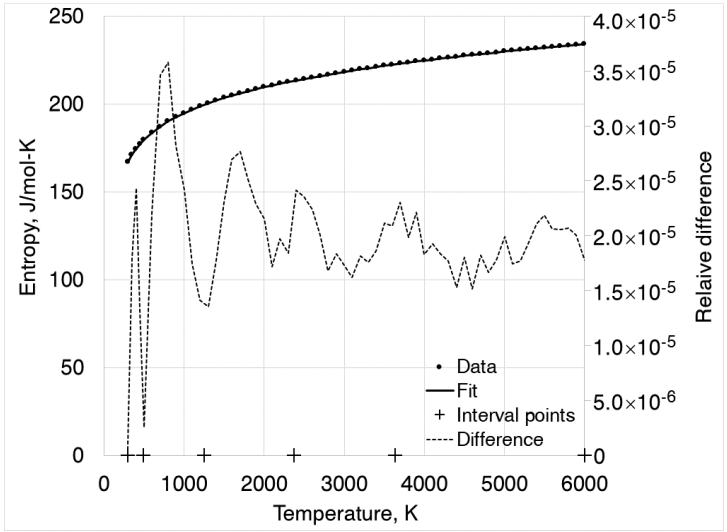


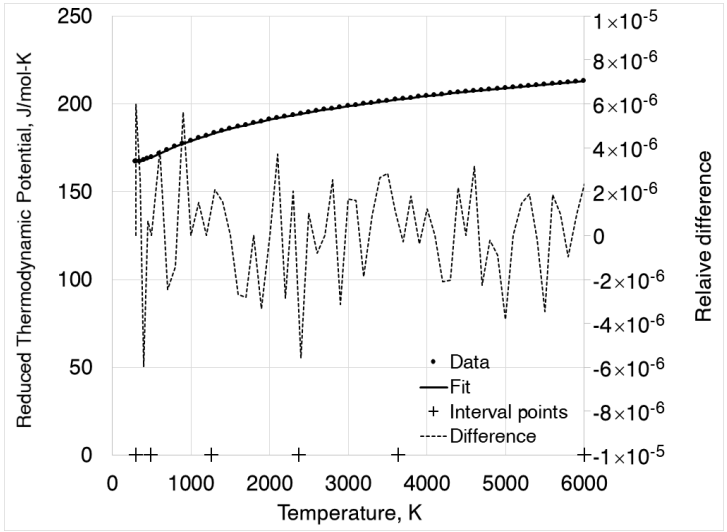
Phosphorus, Ion (P^+) – P1+(g)

MW = 30.97321 g/mol – $\Delta_f H^o$(298.15 K) = 1336.333 kJ/mol – S^o(298.15 K) = 166.97 J/mol-K – p^o = 0.1 MPa



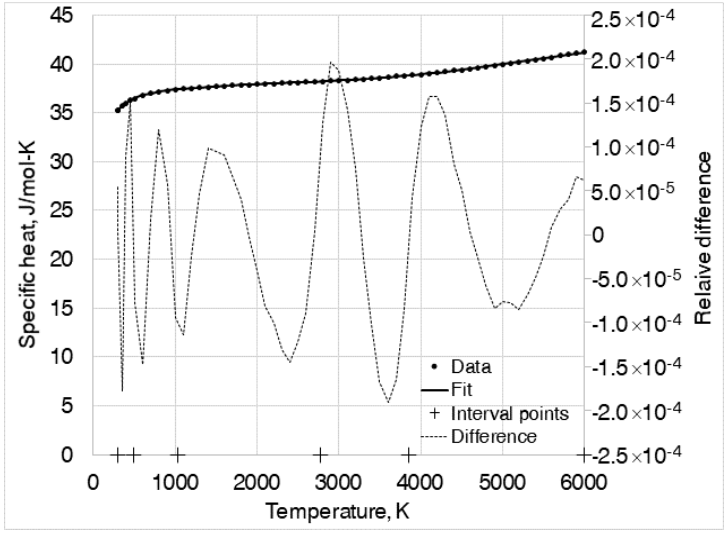


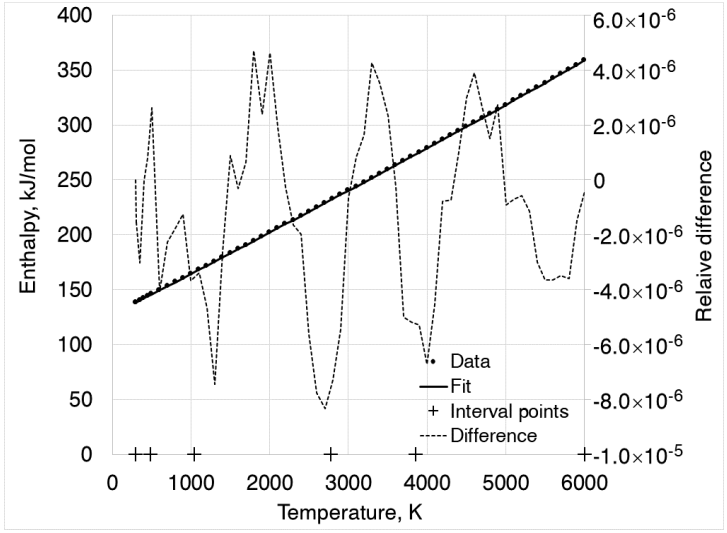


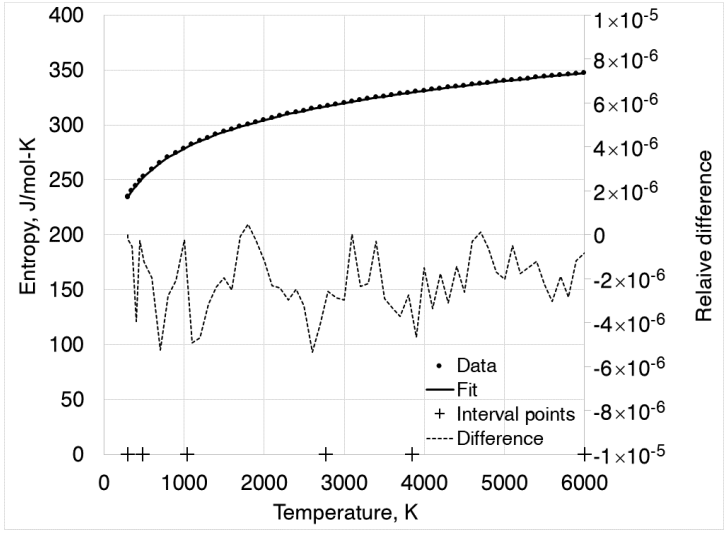


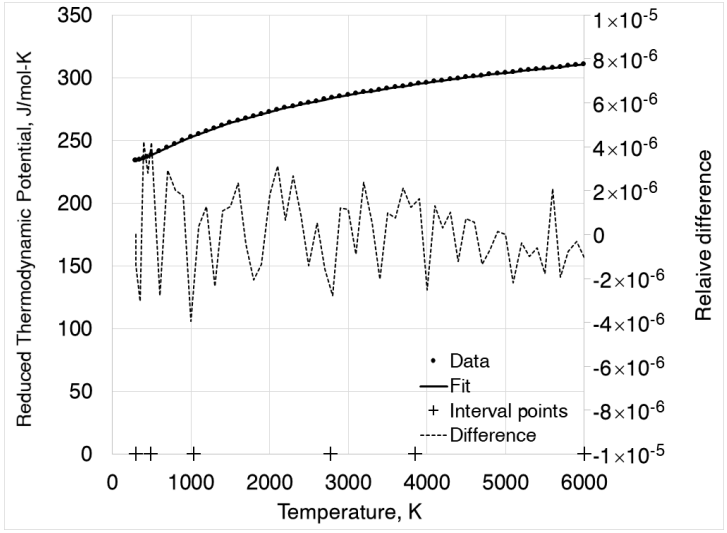
Phosphorus Sulfide (PS) – P1S1(g)

MW = 63.03 g/mol – $\Delta_f H^o$(298.15 K) = 138.603 kJ/mol – S^o(298.15 K) = 234.241 J/mol-K – p^o = 0.1 MPa



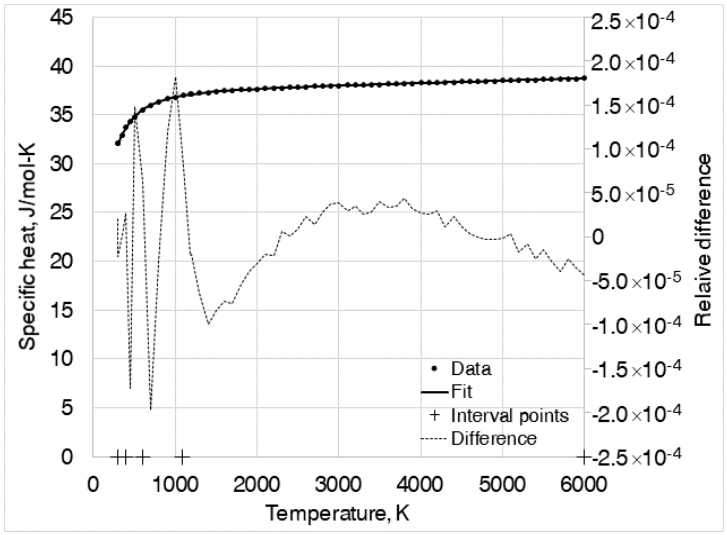


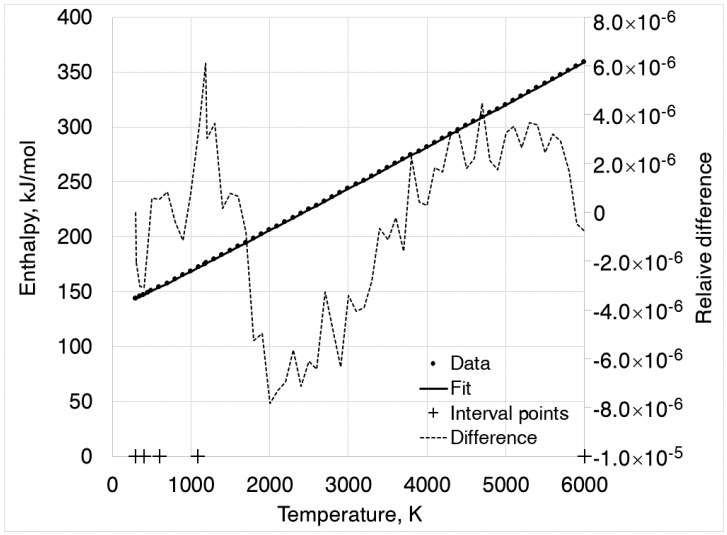


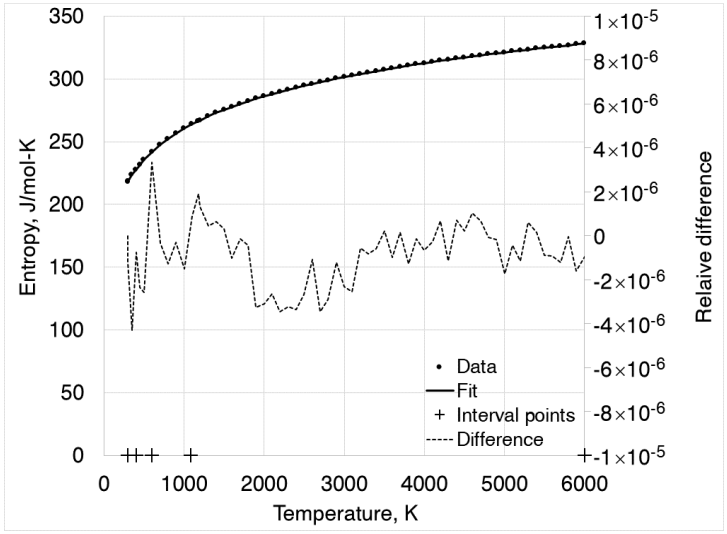


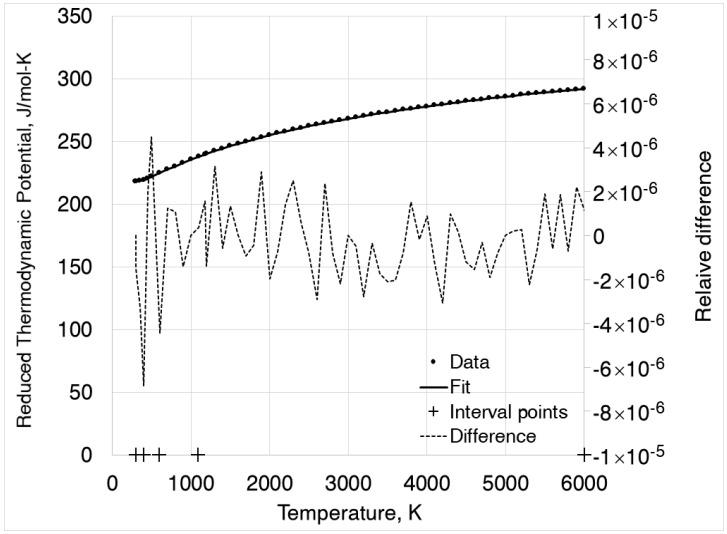
Phosphorus (P_2) – P2(g)

MW = 61.94752 g/mol – $\Delta_f H^o$(298.15 K) = 143.653 kJ/mol – S^o(298.15 K) = 218.131 J/mol-K – p^o = 0.1 MPa



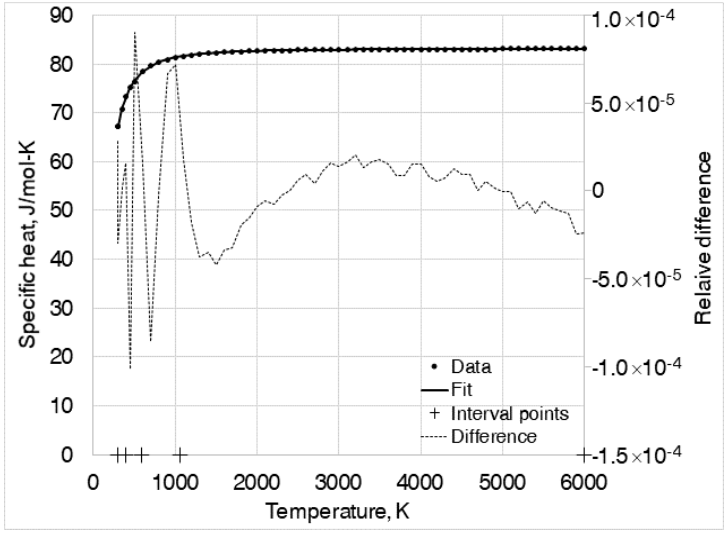


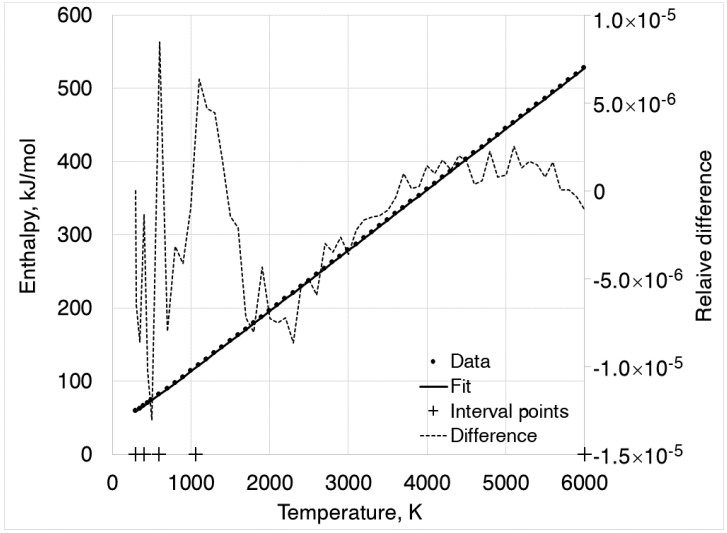


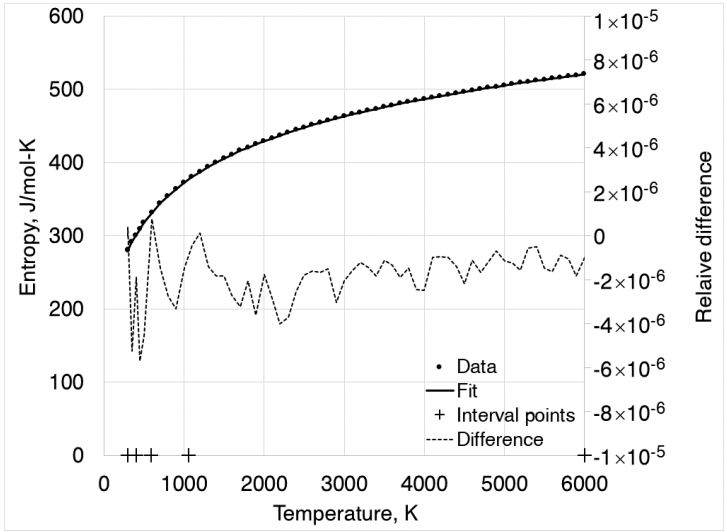


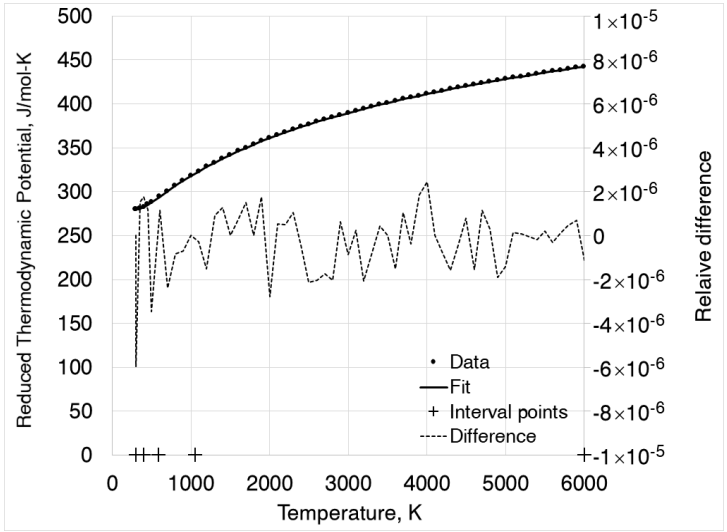
Phosphorus (P_4) – P4(g)

MW = 123.89504 g/mol – $\Delta_f H^o$(298.15 K) = 58.907 kJ/mol – S^o(298.15 K) = 279.992 J/mol-K – p^o = 0.1 MPa



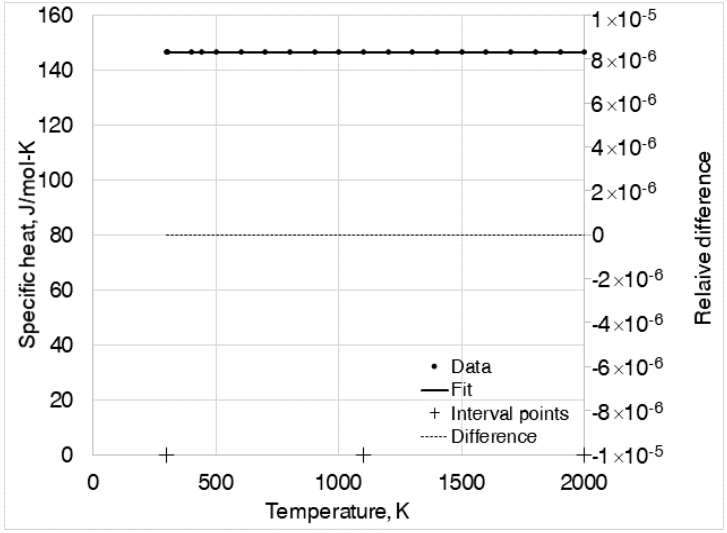


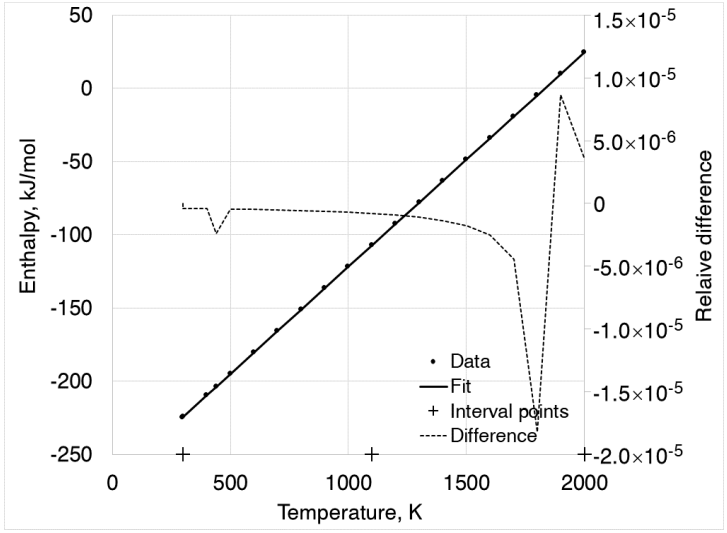


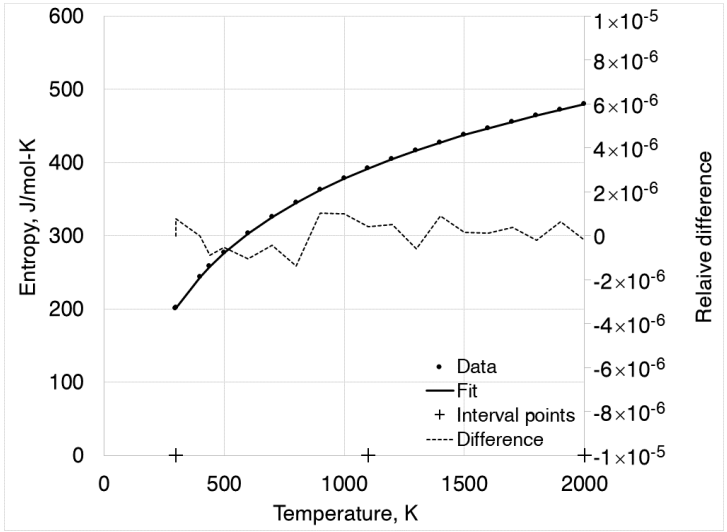


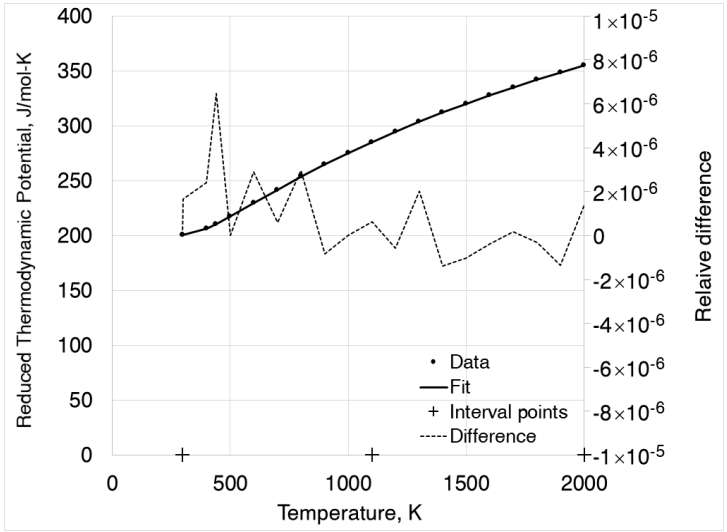
Phosphorus Sulfide (P_4S_3) – P4S3(cr)

MW = 220.08 g/mol – $\Delta_f H^o$(298.15 K) = -224.647 kJ/mol – S^o(298.15 K) = 200.832 J/mol-K – p^o = 0.1 MPa



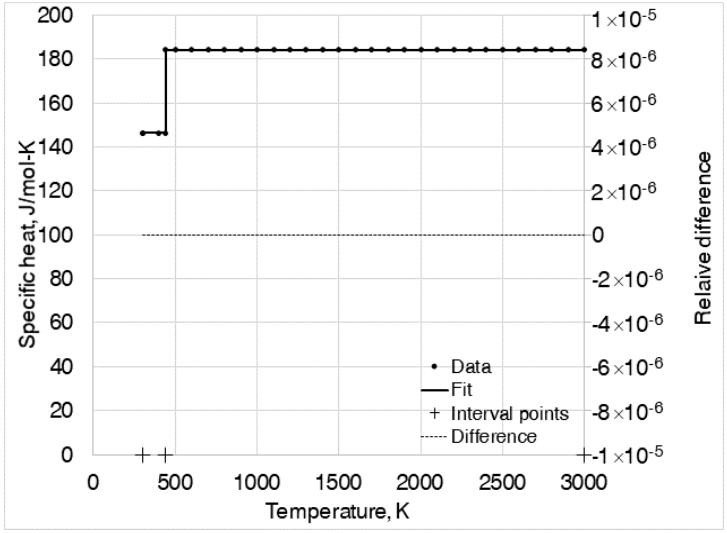


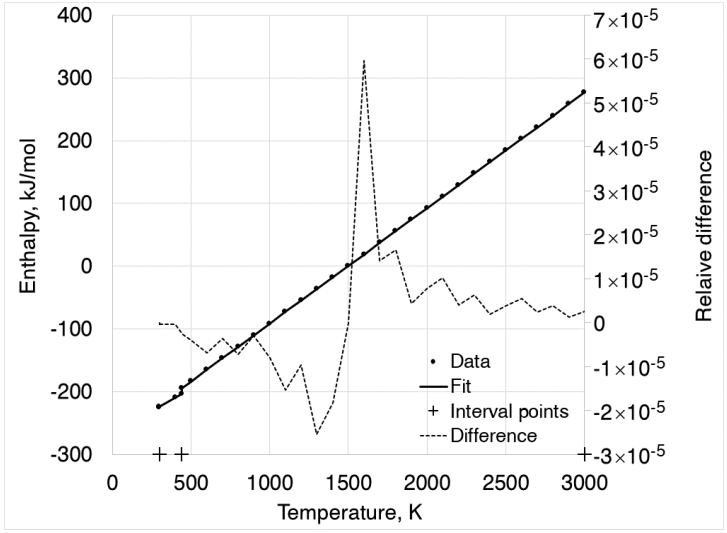


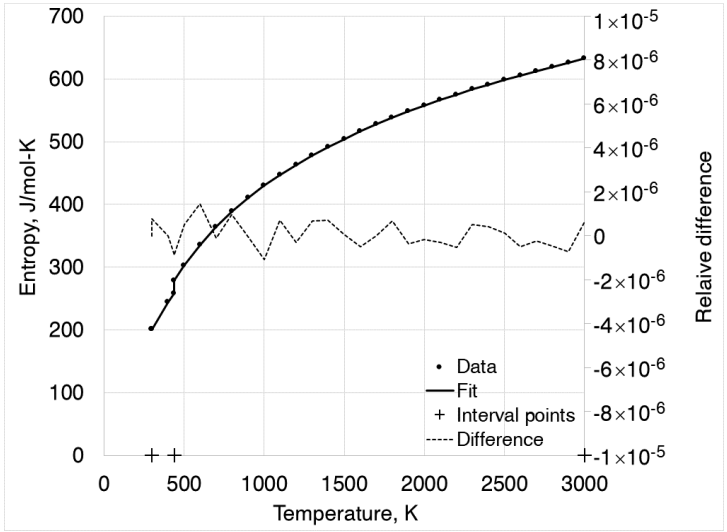


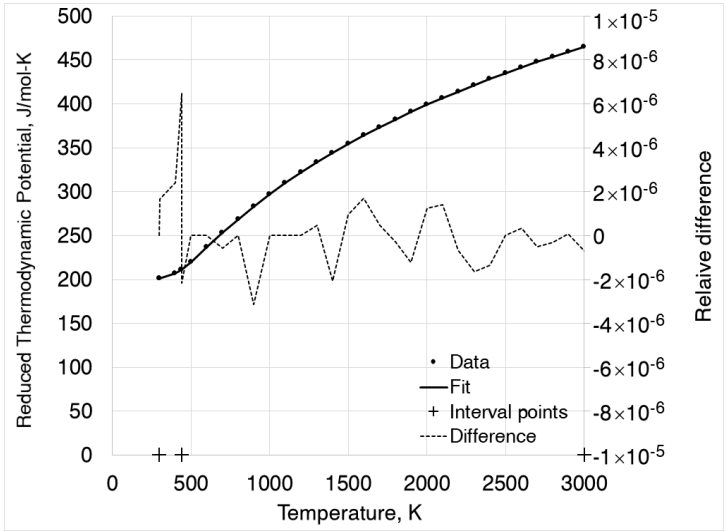
Phosphorus Sulfide (P_4S_3) – P4S3(cr,l)

MW = 220.08 g/mol – $\Delta_f H^o$(298.15 K) = -224.647 kJ/mol – S^o(298.15 K) = 200.832 J/mol-K – p^o = 0.1 MPa



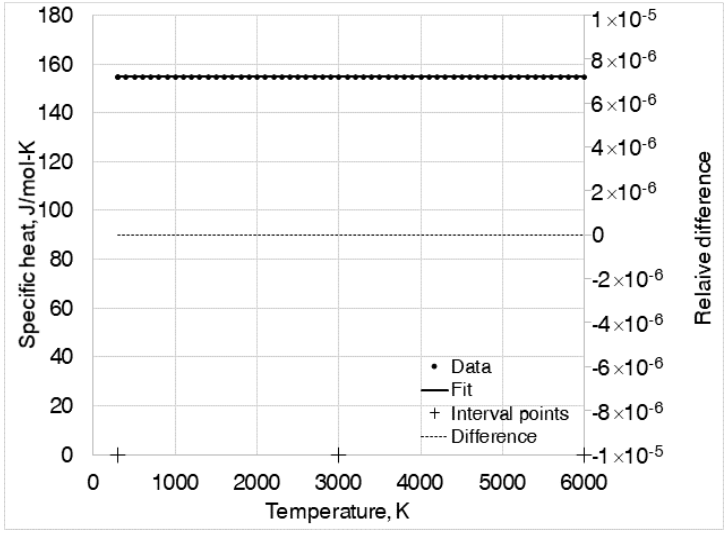


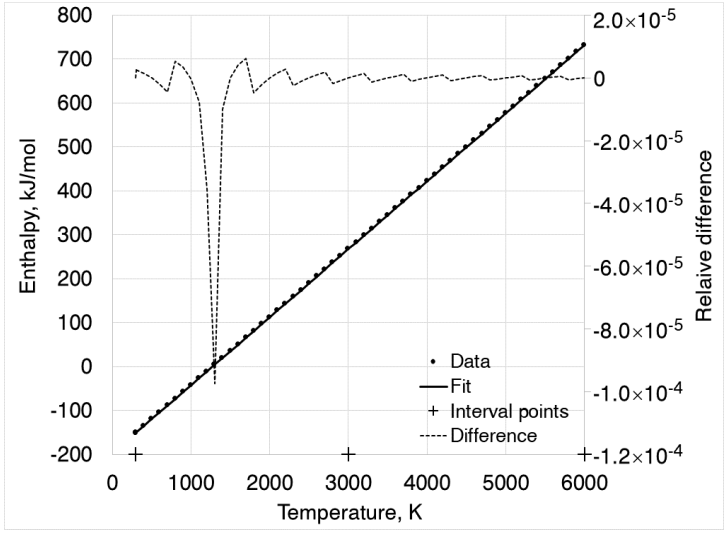


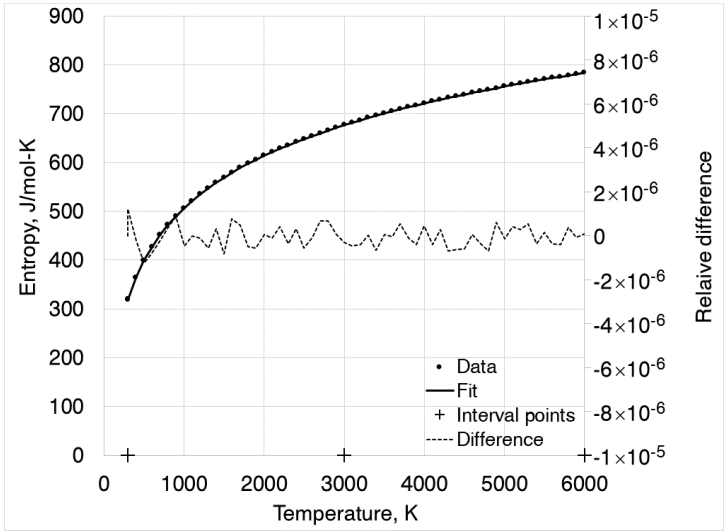


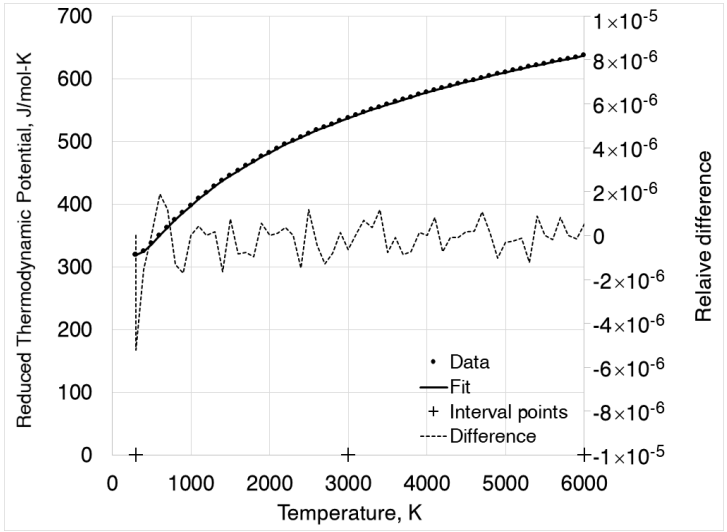
Phosphorus Sulfide (P_4S_3) – P4S3(g)

MW = 220.08 g/mol – $\Delta_f H^o$(298.15 K) = -151.042 kJ/mol – S^o(298.15 K) = 319.265 J/mol-K – p^o = 0.1 MPa



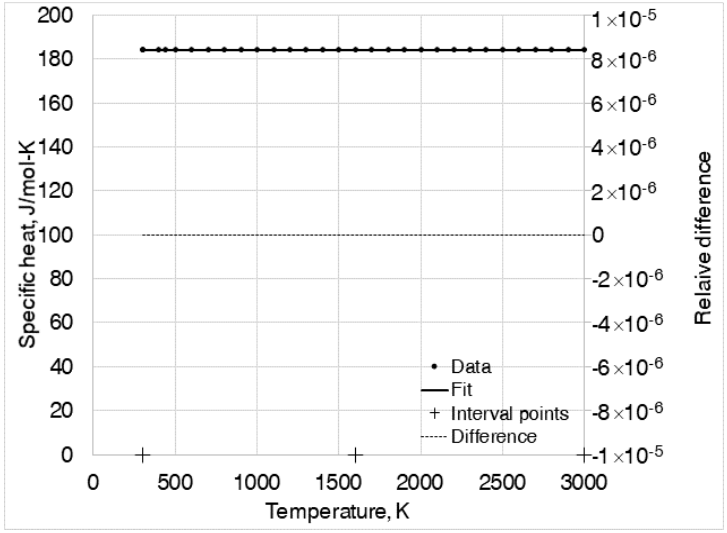


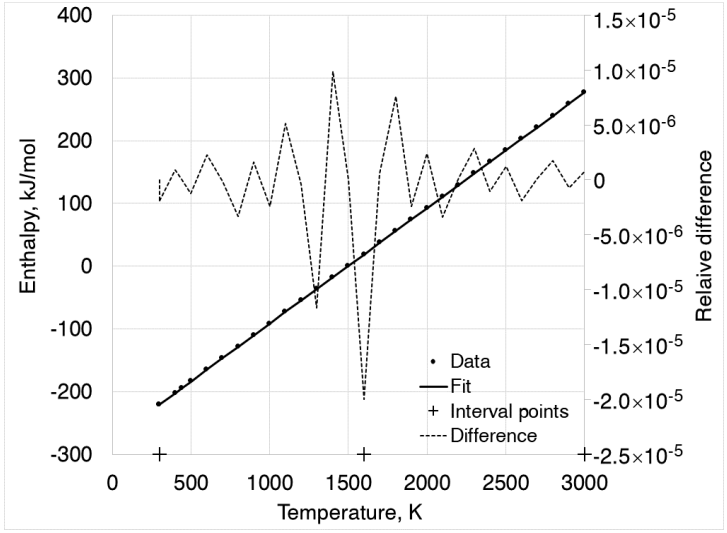


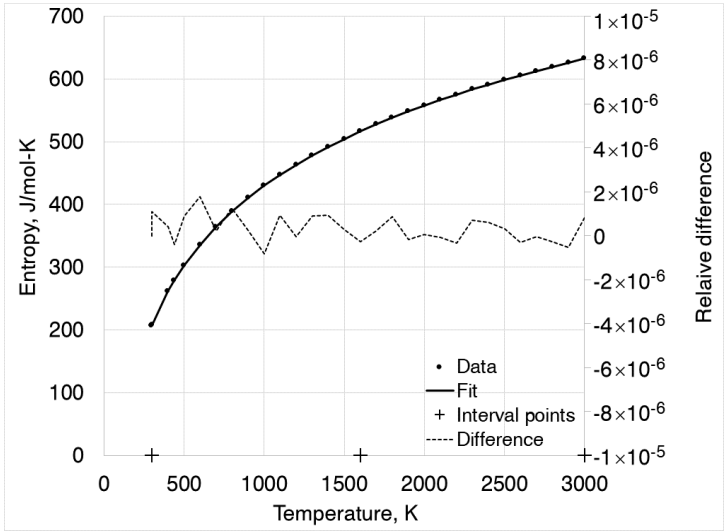


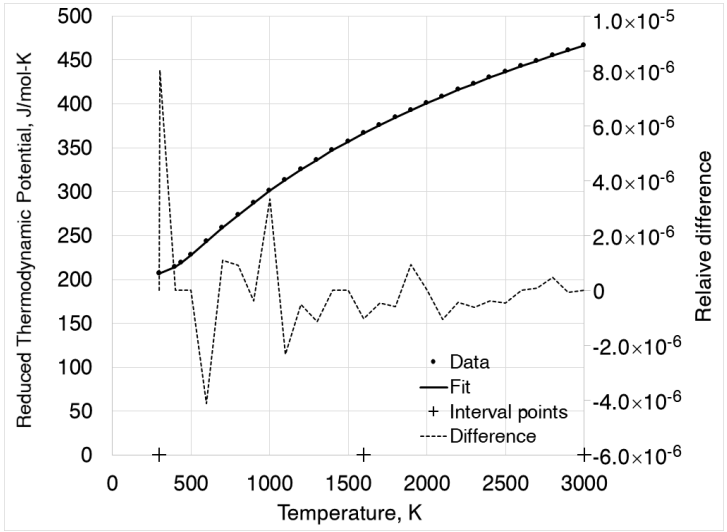
Phosphorus Sulfide (P_4S_3) – P4S3(l)

MW = 220.08 g/mol – $\Delta_f H^o$(298.15 K) = -220.784 kJ/mol – S^o(298.15 K) = 207.097 J/mol-K – p^o = 0.1 MPa



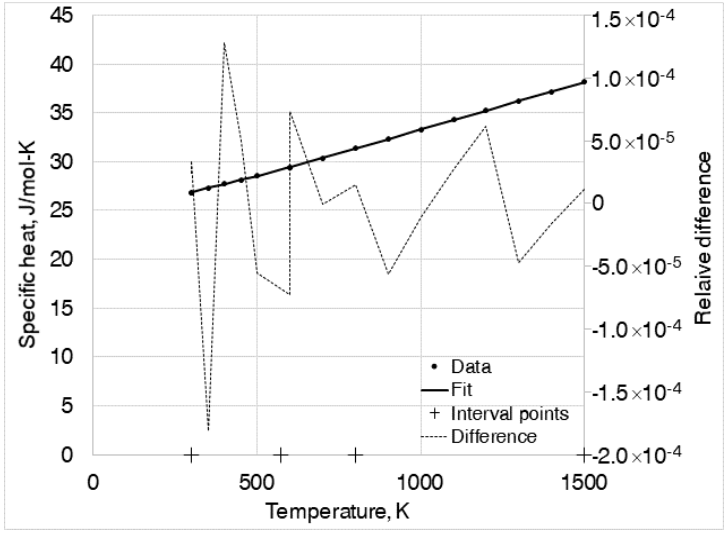


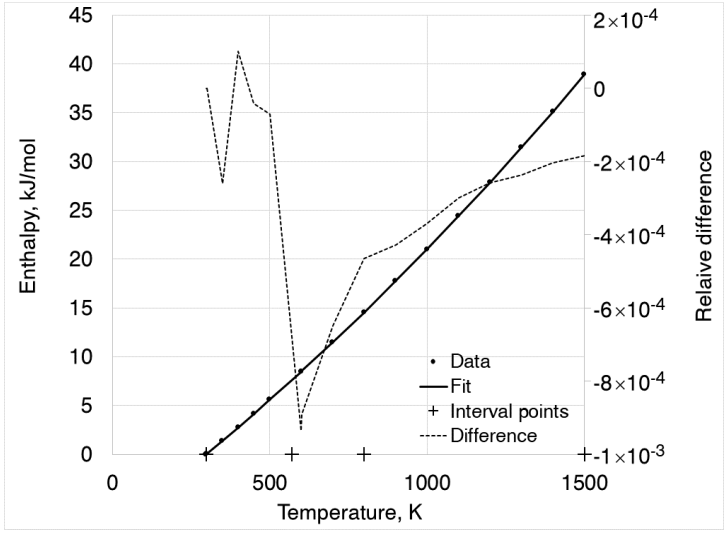


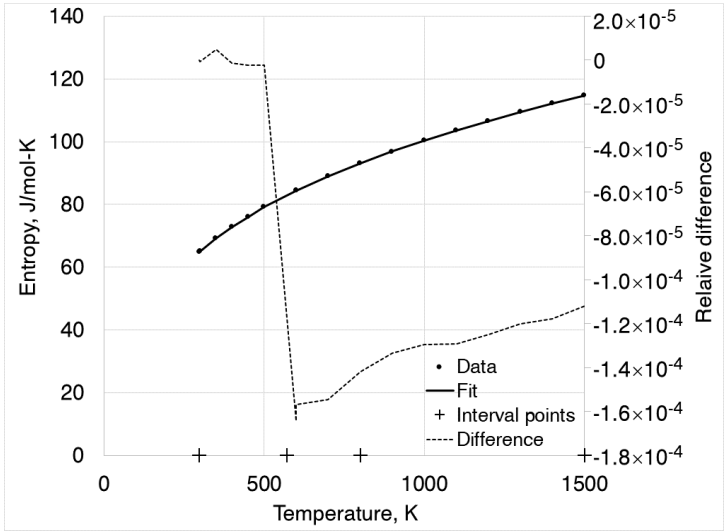


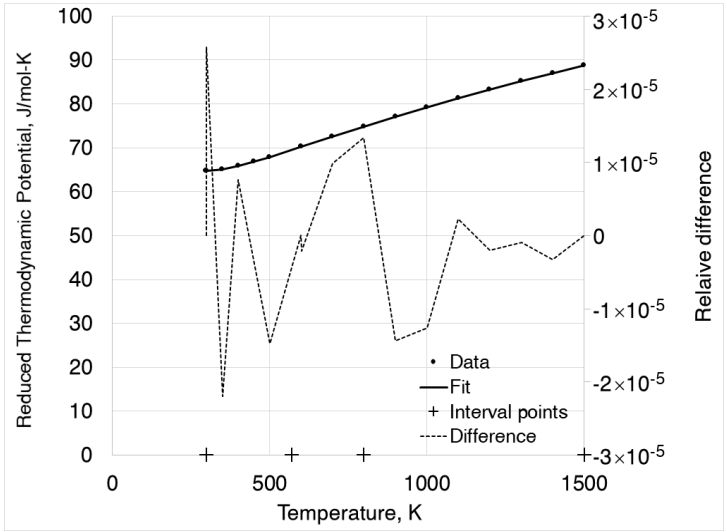
Lead (Pb) – Pb1(cr)

MW = 207.2 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 64.785 J/mol-K – p^o = 0.1 MPa



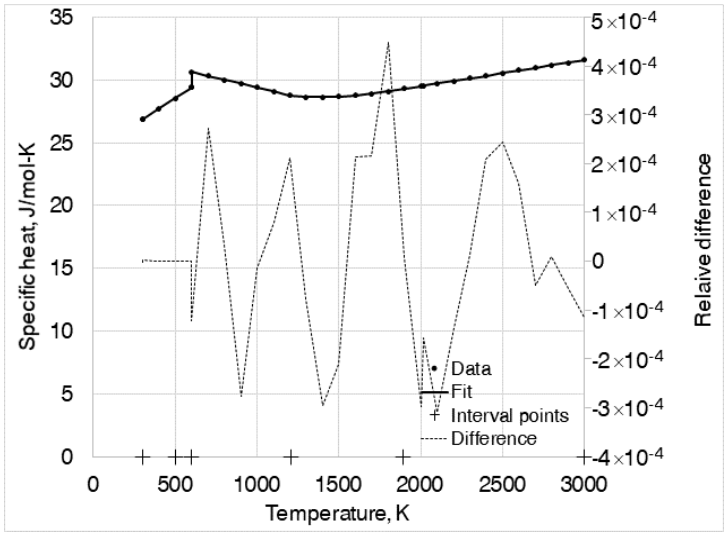


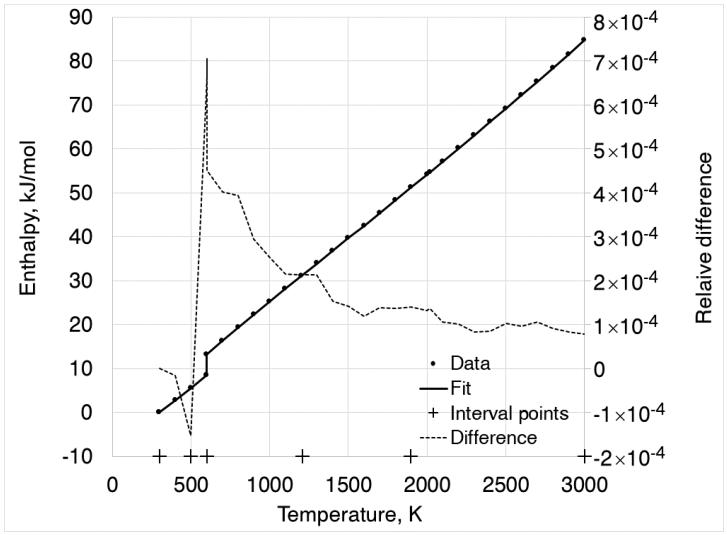


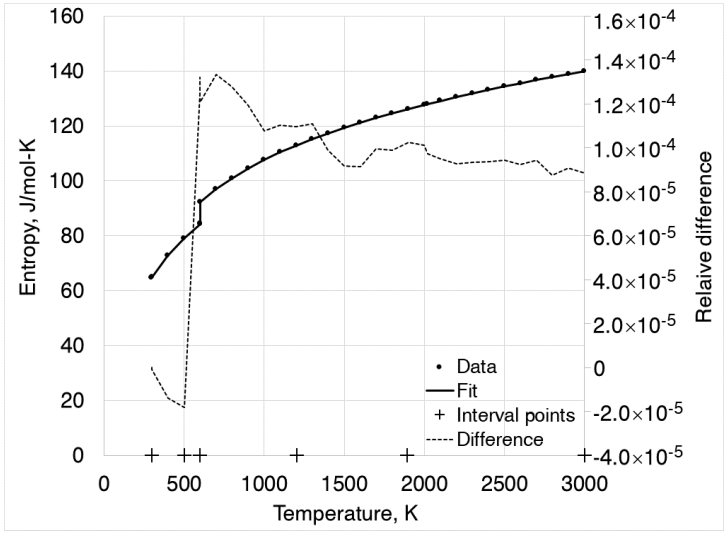


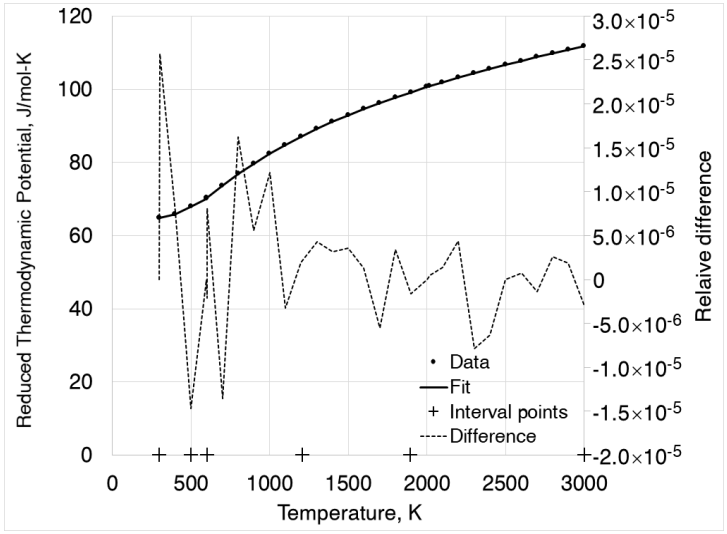
Lead (Pb) – Pb1(cr,l)

MW = 207.2 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 64.785 J/mol-K – p^o = 0.1 MPa



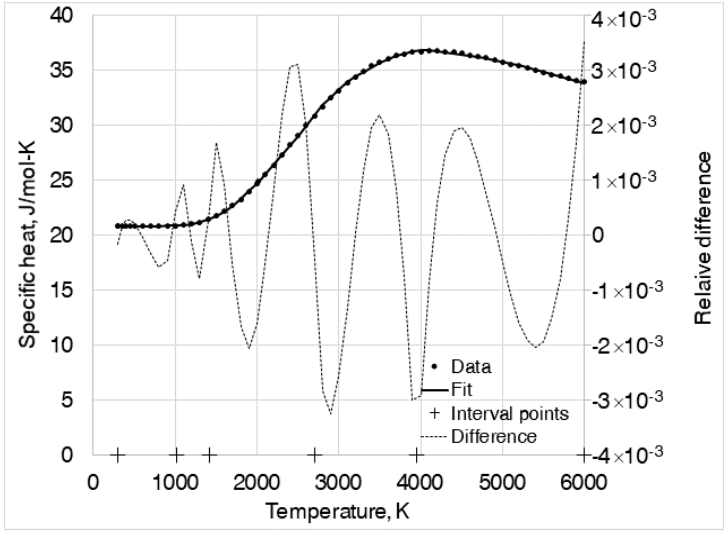


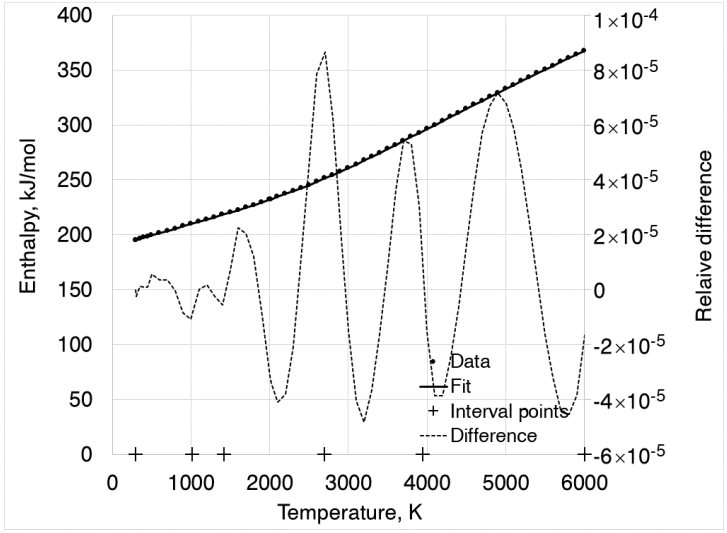


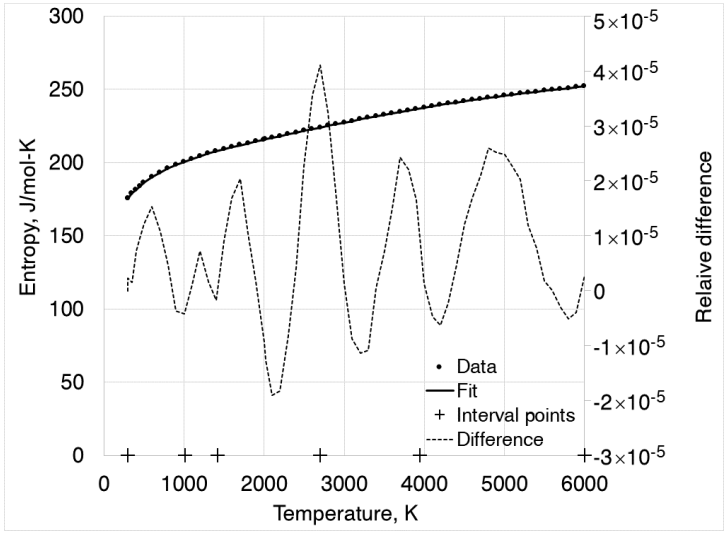


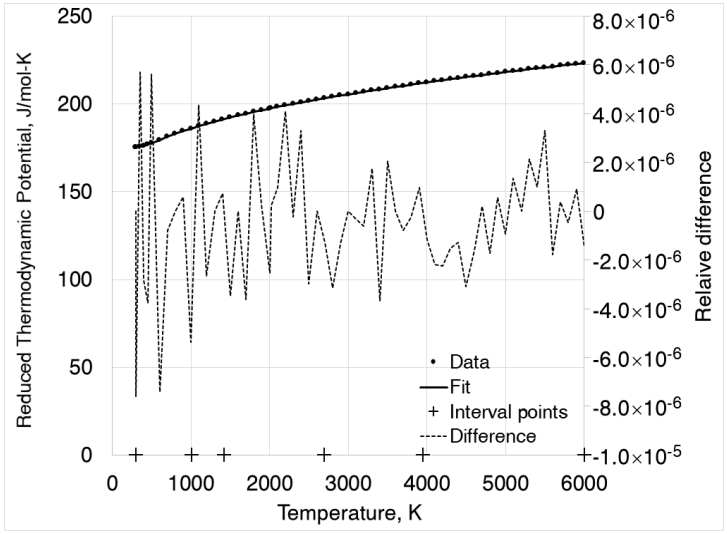
Lead (Pb) – Pb1(g)

MW = 207.2 g/mol – $\Delta_f H^o$(298.15 K) = 195.2 kJ/mol – S^o(298.15 K) = 175.374 J/mol-K – p^o = 0.1 MPa



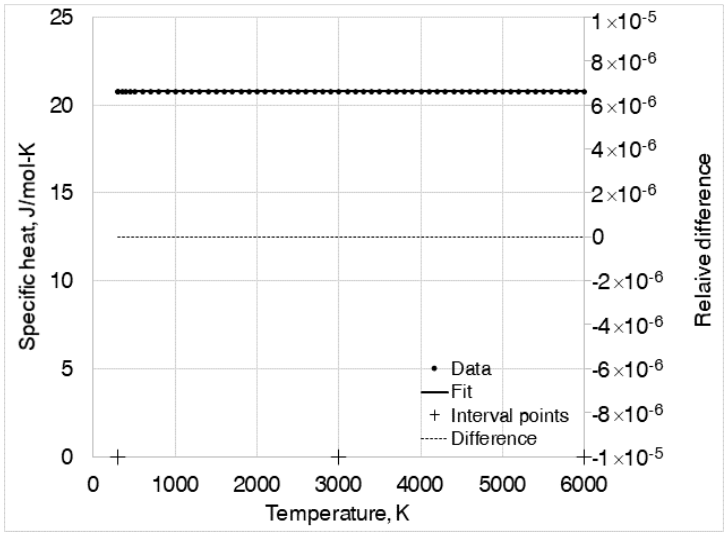


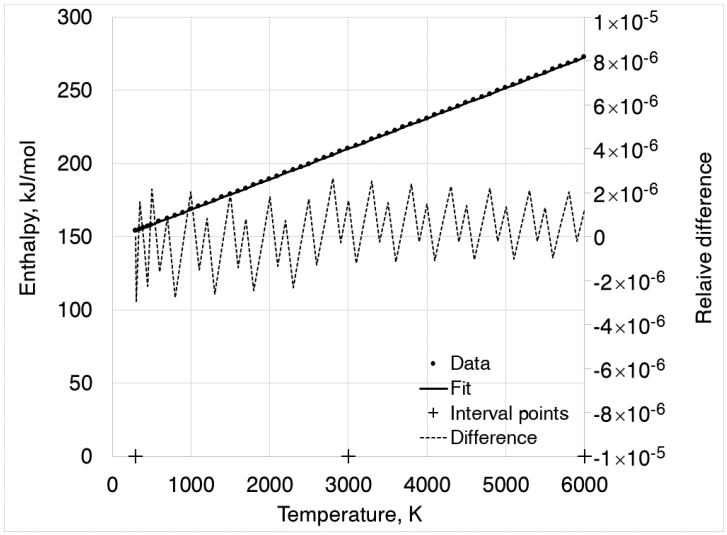


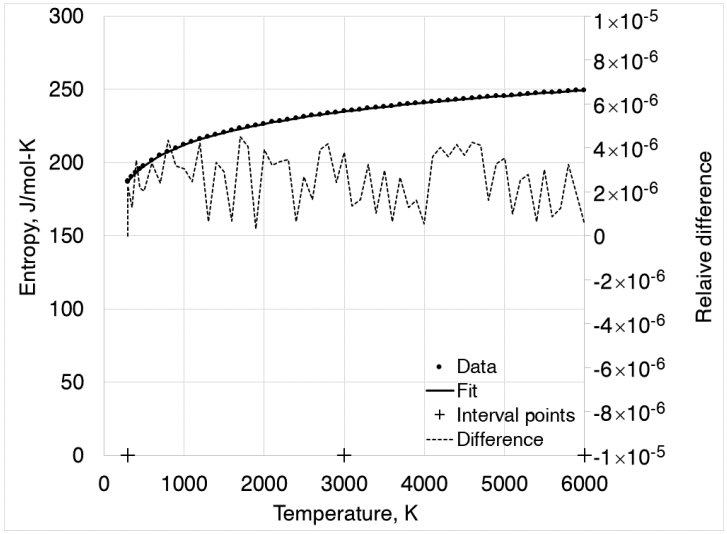


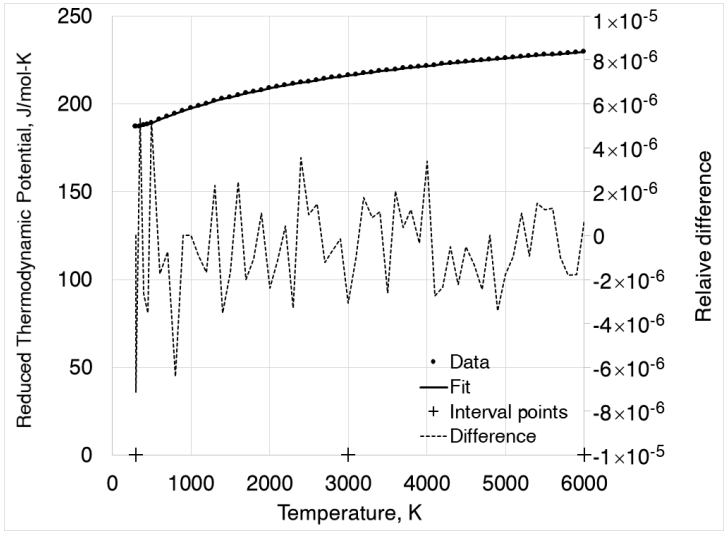
Lead, Ion (Pb^-) – Pb1-(g)

MW = 207.2 g/mol – $\Delta_f H^o$(298.15 K) = 153.882 kJ/mol – S^o(298.15 K) = 186.9 J/mol-K – p^o = 0.1 MPa



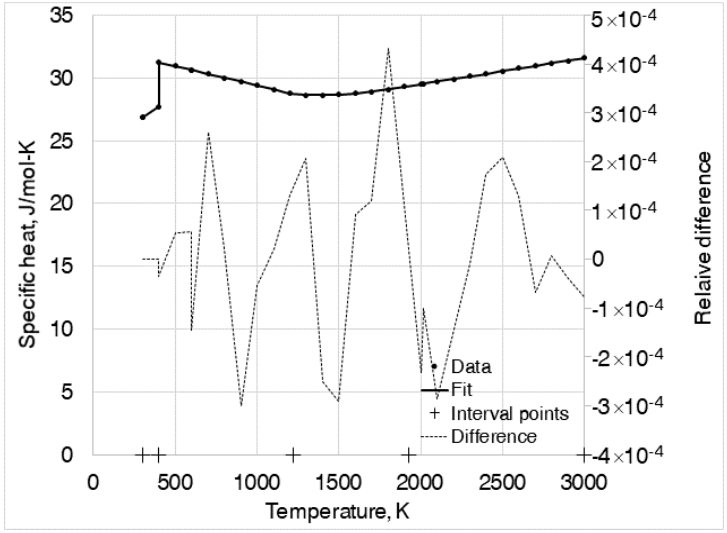


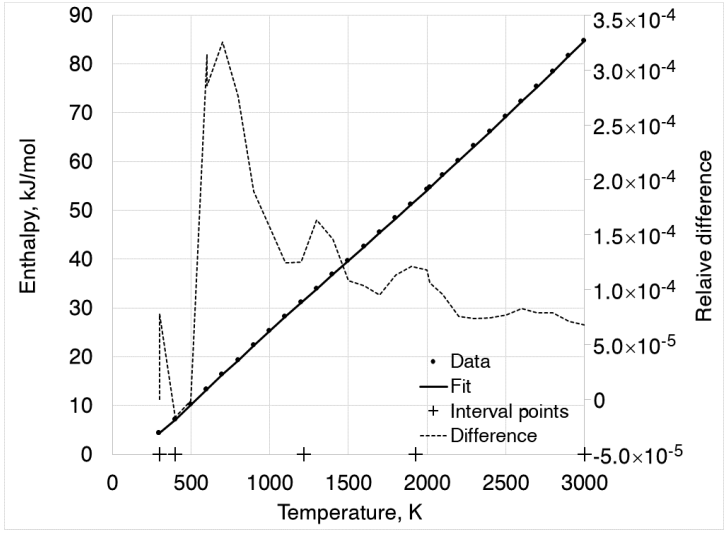


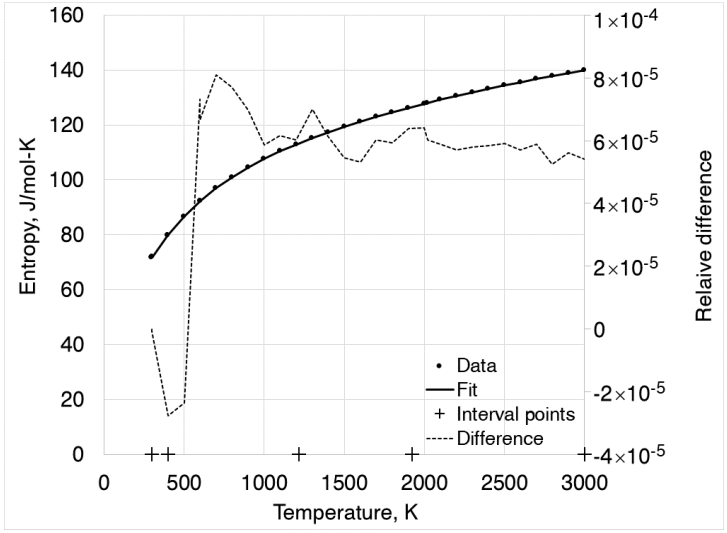


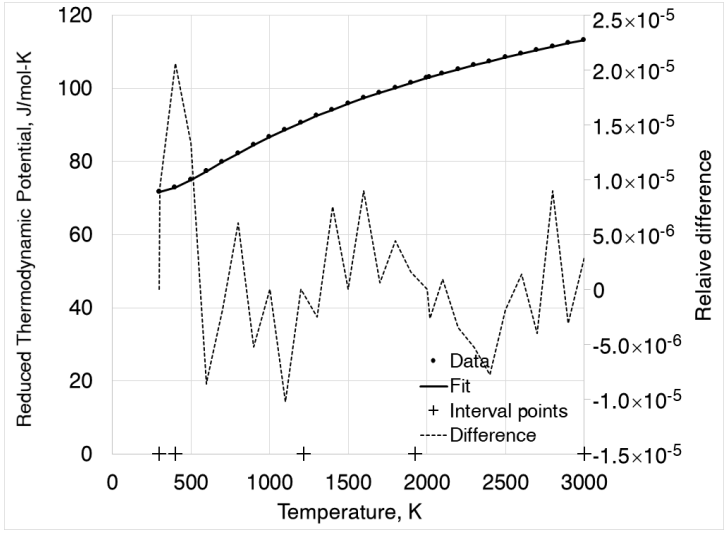
Lead (Pb) – Pb1(l)

MW = 207.2 g/mol – $\Delta_f H^o$(298.15 K) = 4.283 kJ/mol – S^o(298.15 K) = 71.71 J/mol-K – p^o = 0.1 MPa



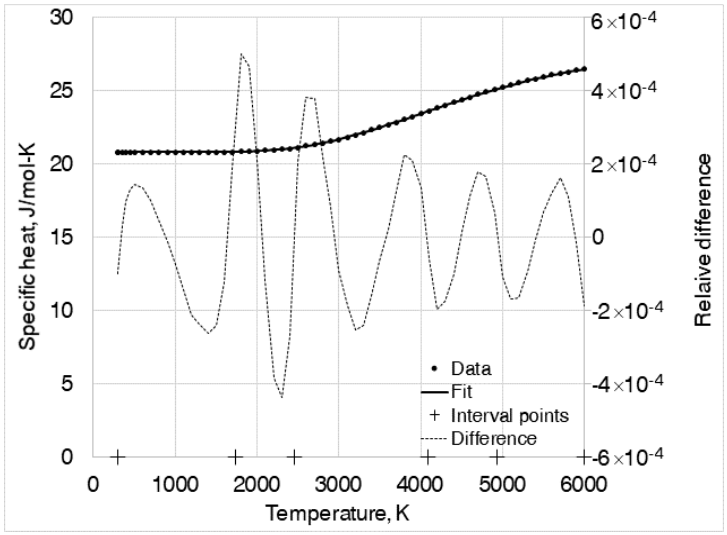


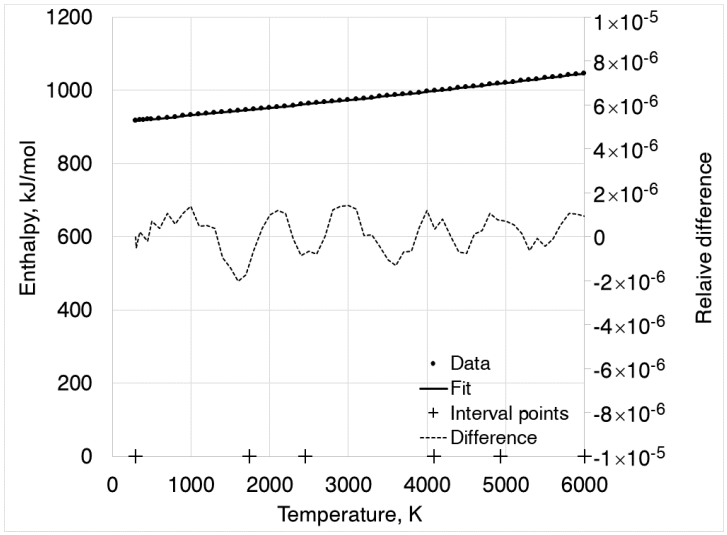


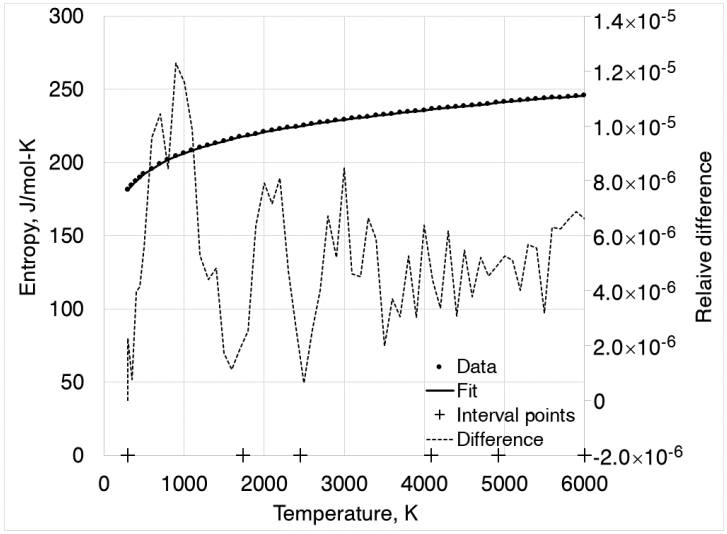


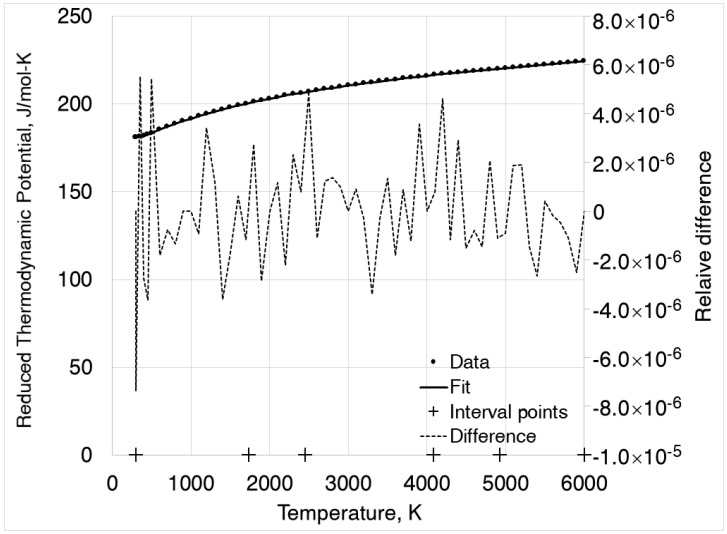
Lead, Ion (Pb^+) – Pb1+(g)

MW = 207.2 g/mol – $\Delta_f H^o$(298.15 K) = 916.996 kJ/mol – S^o(298.15 K) = 181.137 J/mol-K – p^o = 0.1 MPa



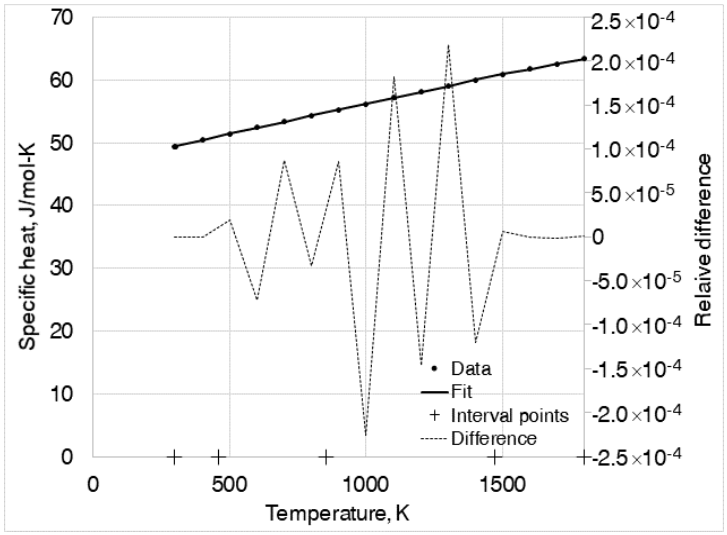


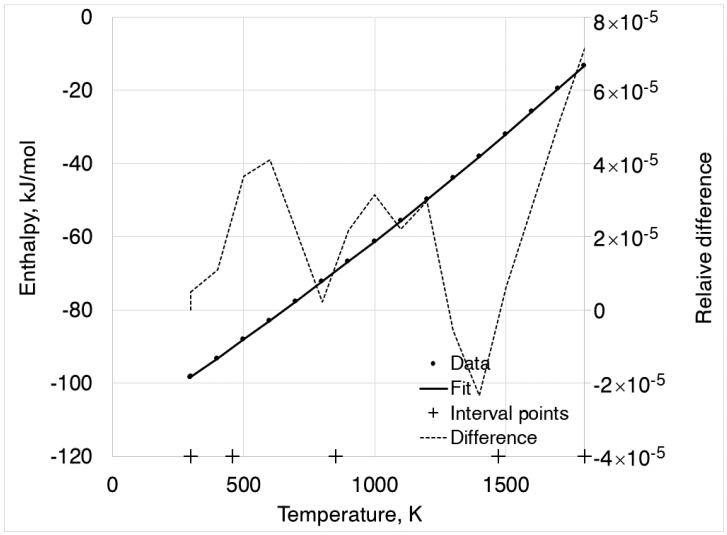


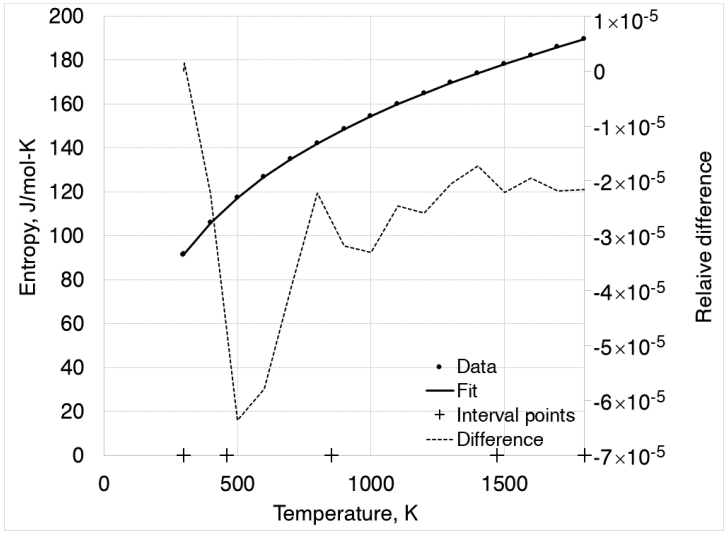


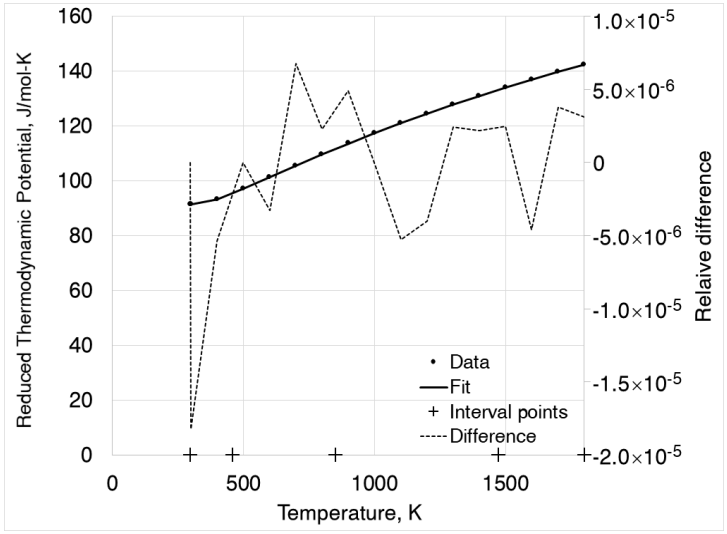
Lead Sulfide (PbS) – Pb1S1(cr)

MW = 239.3 g/mol – $\Delta_f H^o$(298.15 K) = -98.324 kJ/mol – S^o(298.15 K) = 91.343 J/mol-K – p^o = 0.1 MPa



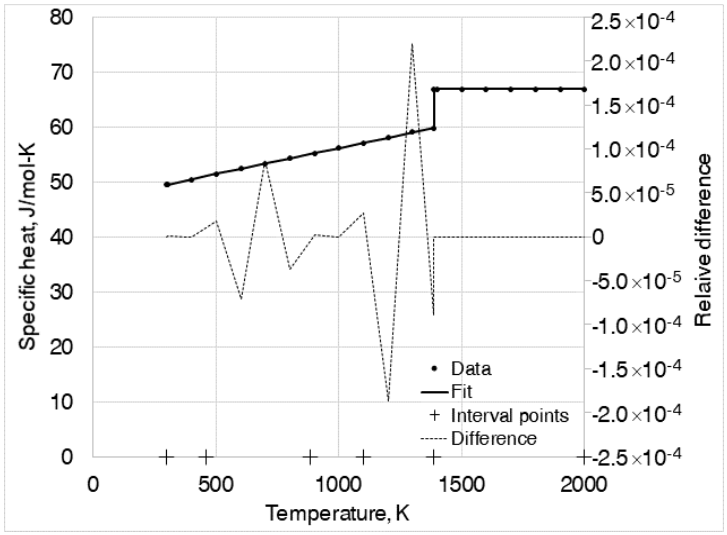


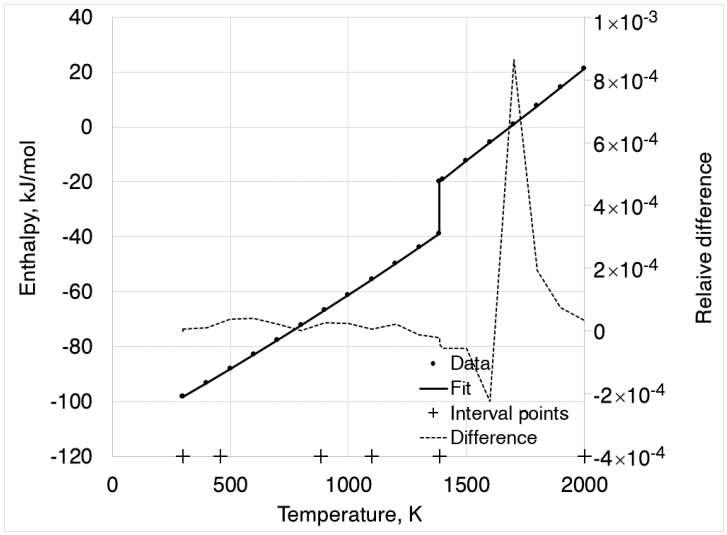


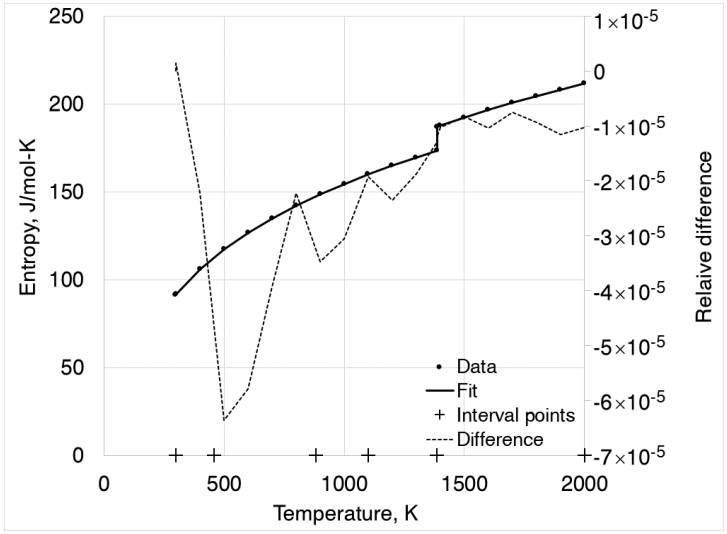


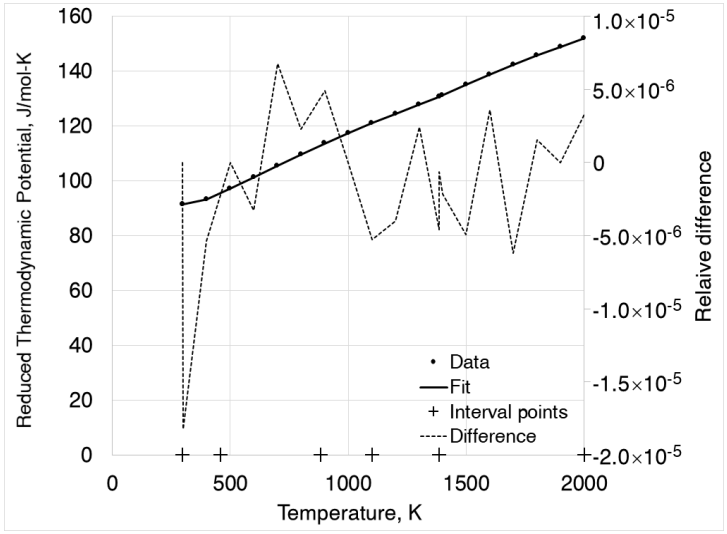
Lead Sulfide (PbS) – Pb1S1(cr,l)

MW = 239.3 g/mol – $\Delta_f H^o$(298.15 K) = -98.324 kJ/mol – S^o(298.15 K) = 91.343 J/mol-K – p^o = 0.1 MPa



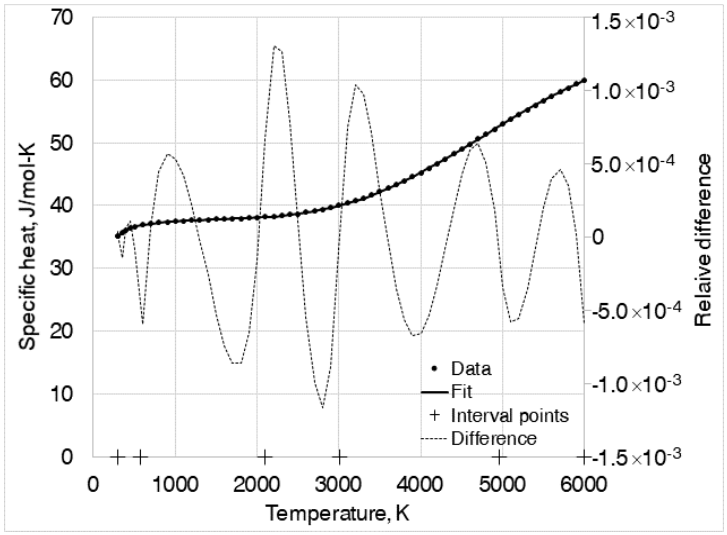


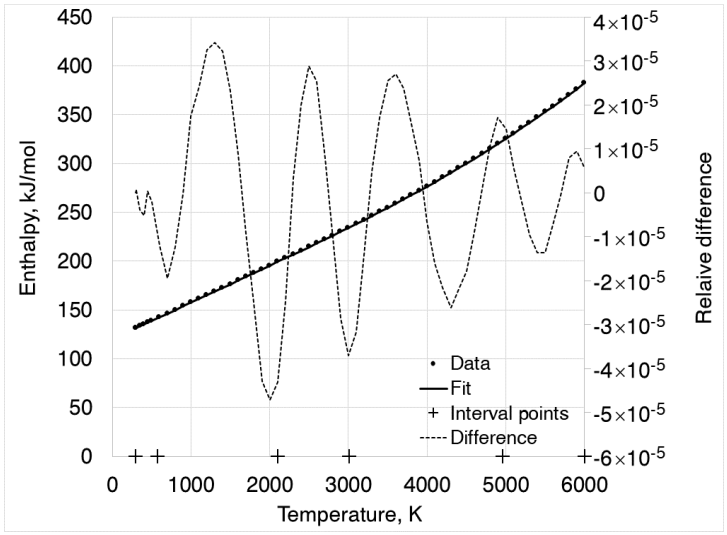


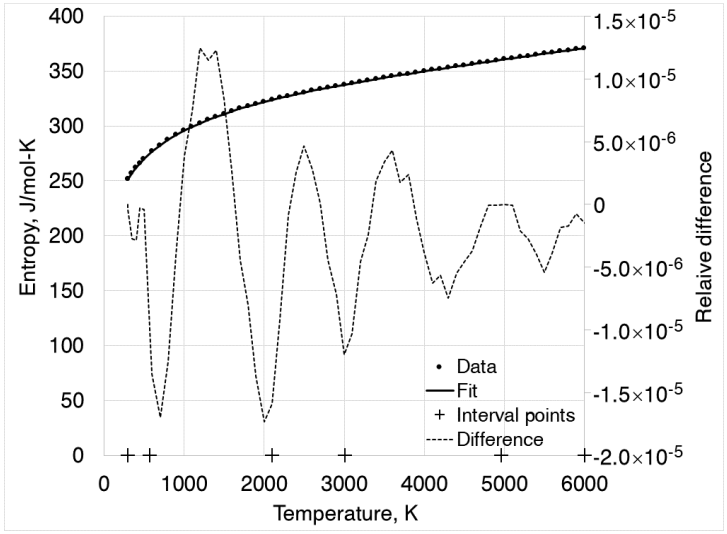


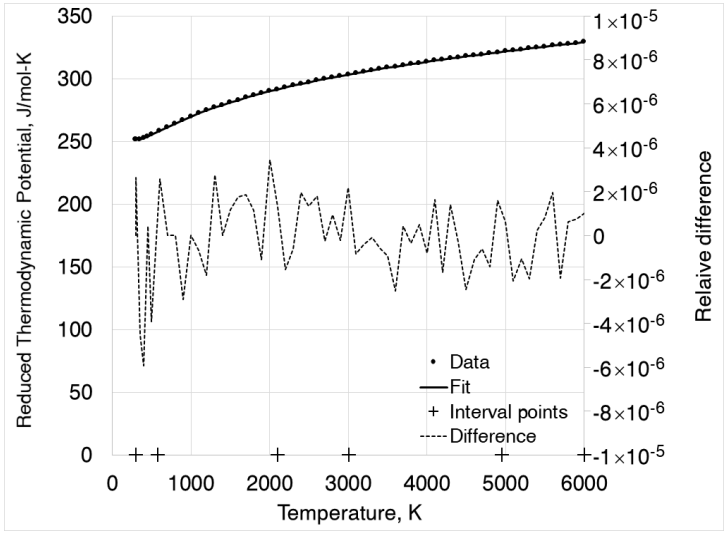
Lead Sulfide (PbS) – Pb1S1(g)

MW = 239.3 g/mol – $\Delta_f H^o$(298.15 K) = 131.796 kJ/mol – S^o(298.15 K) = 251.412 J/mol-K – p^o = 0.1 MPa



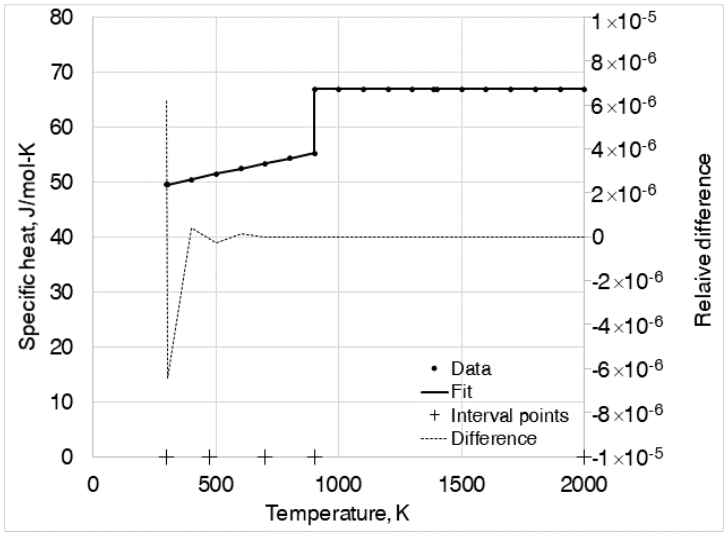


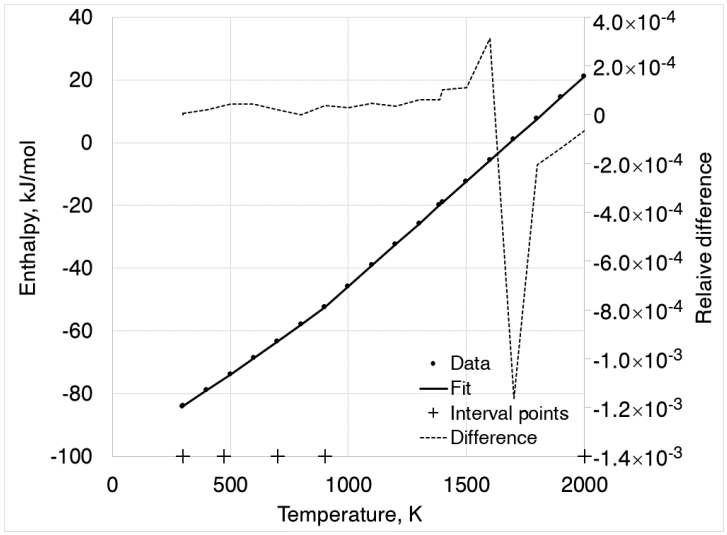


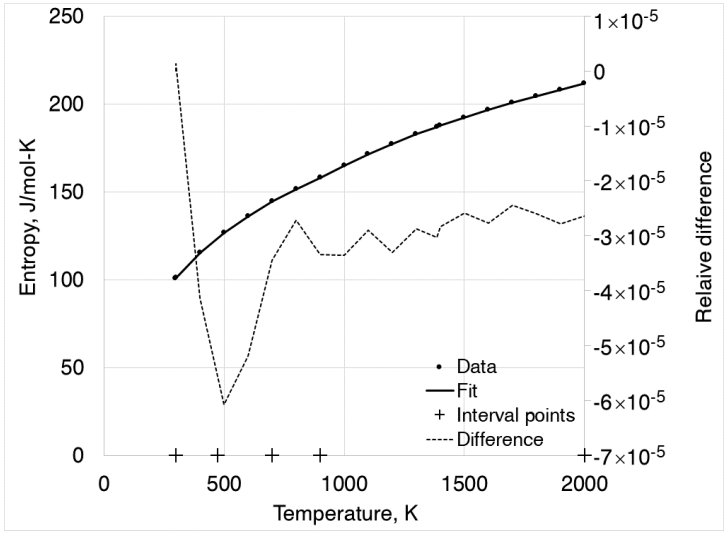


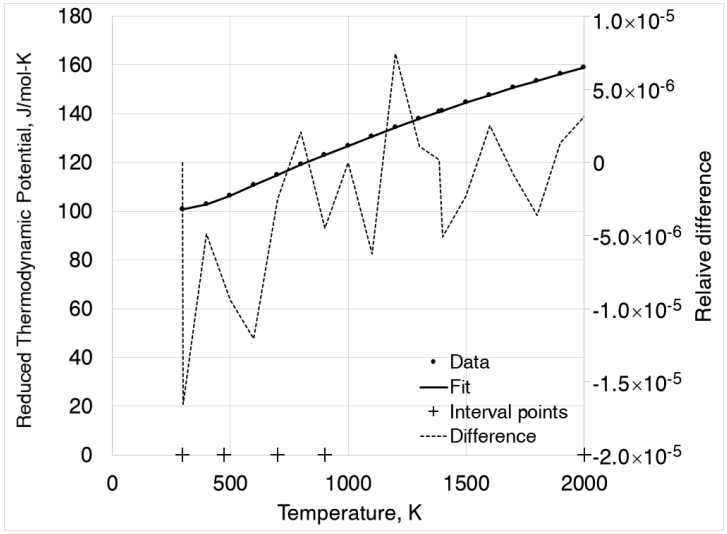
Lead Sulfide (PbS) – Pb1S1(l)

MW = 239.3 g/mol – $\Delta_f H^o$(298.15 K) = -84.064 kJ/mol – S^o(298.15 K) = 100.795 J/mol-K – p^o = 0.1 MPa



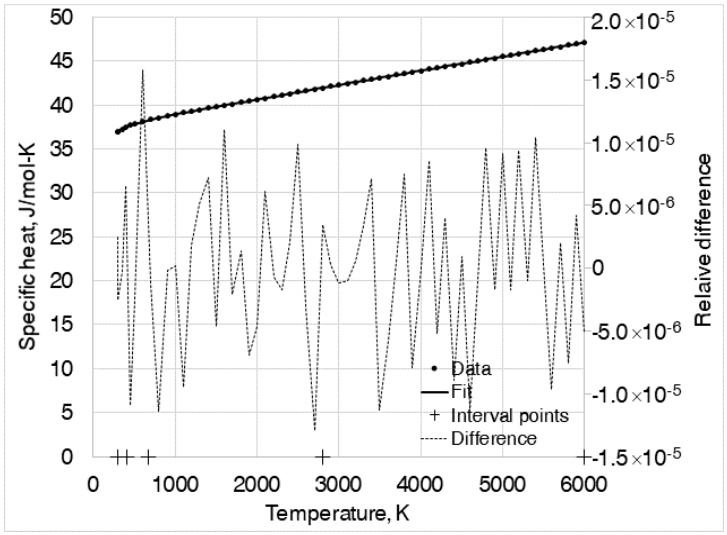


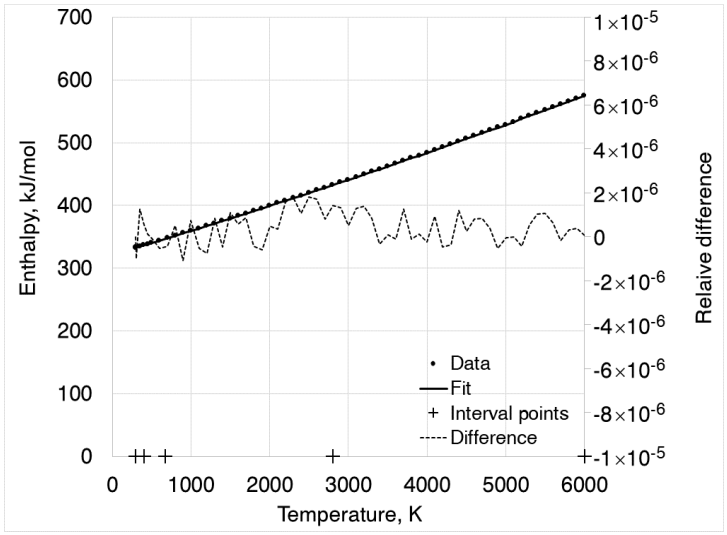


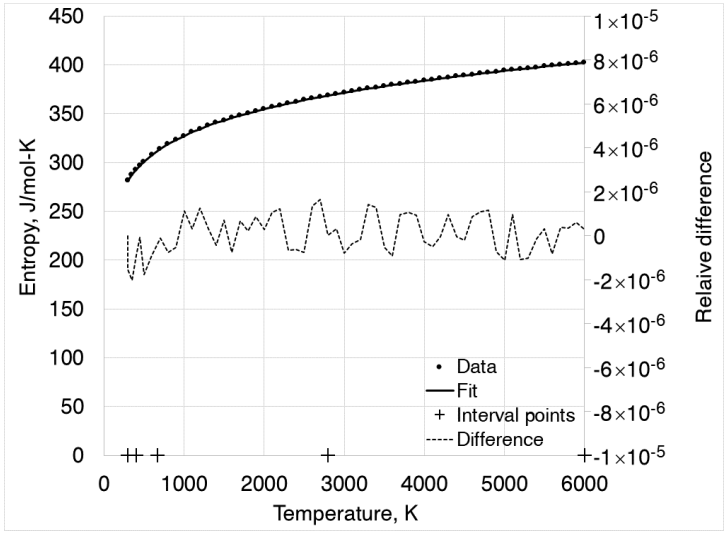


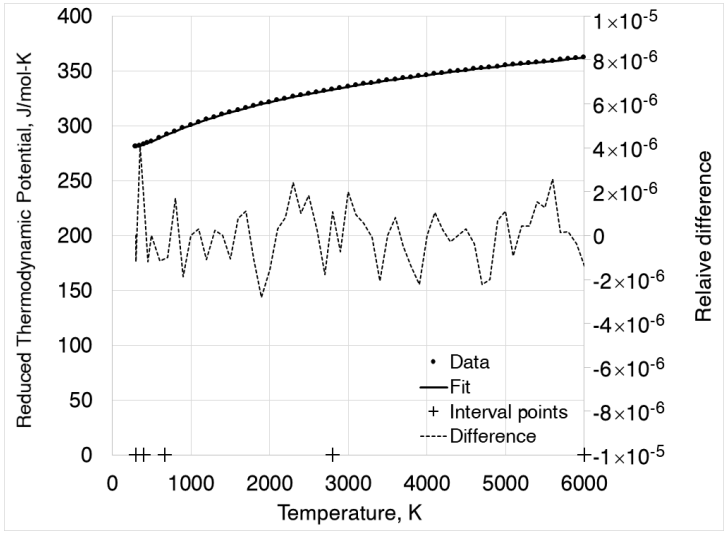
Lead (Pb_2) – Pb2(g)

MW = 414.4 g/mol – $\Delta_f H^o$(298.15 K) = 332.628 kJ/mol – S^o(298.15 K) = 281.334 J/mol-K – p^o = 0.1 MPa



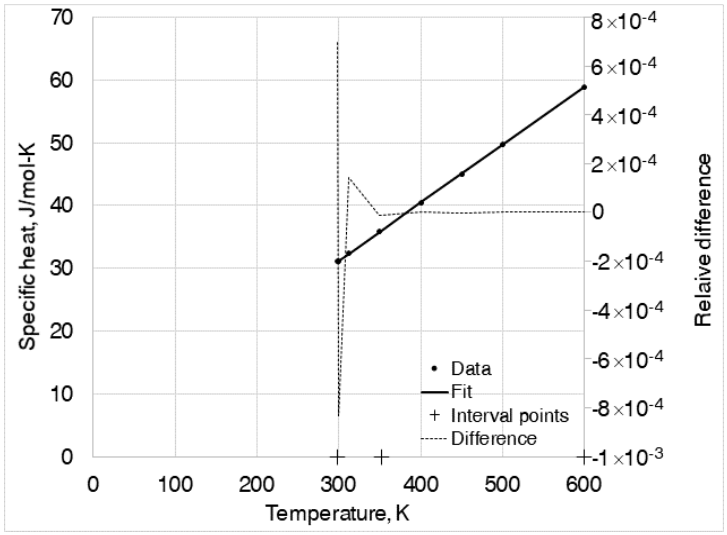


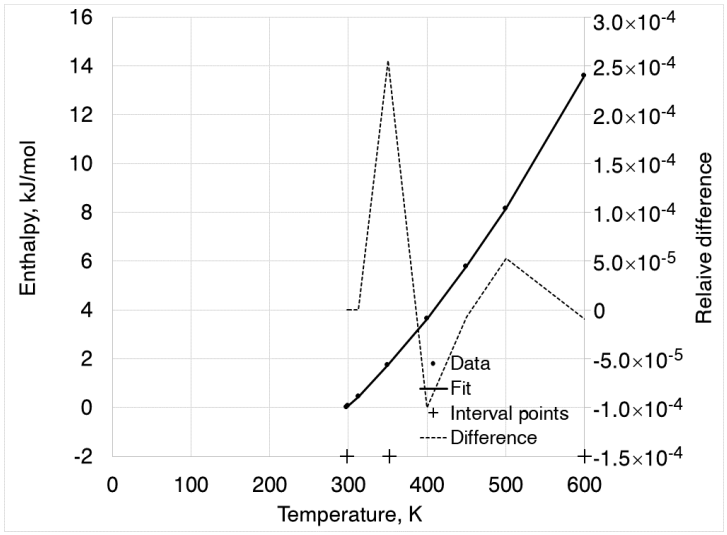


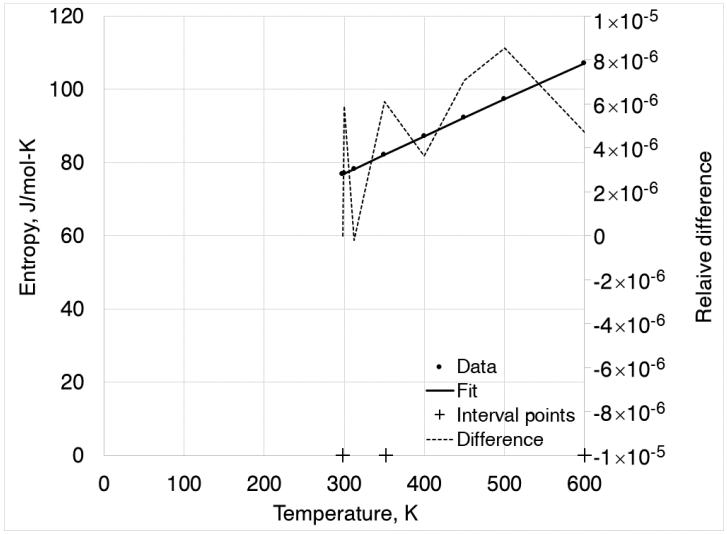


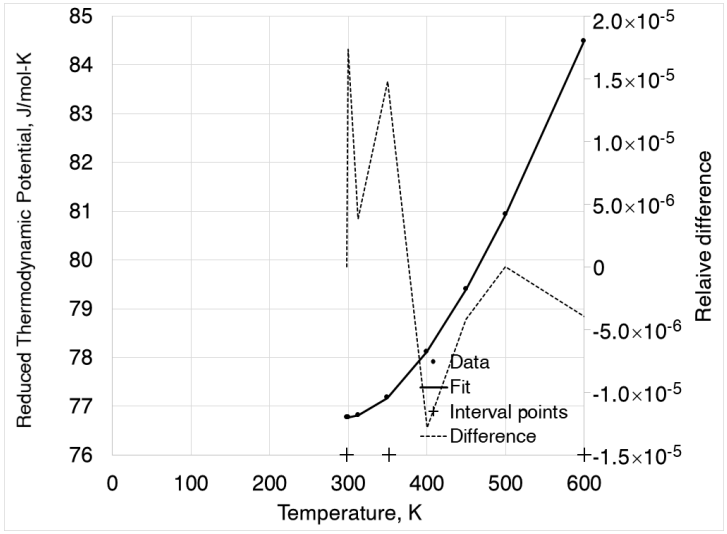
Rubidium (Rb) – Rb1(cr)

MW = 85.4678 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 76.778 J/mol-K – p^o = 0.1 MPa



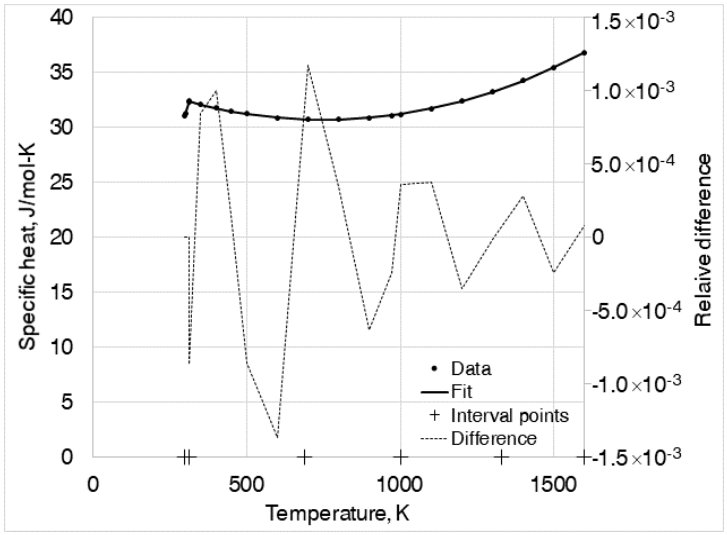


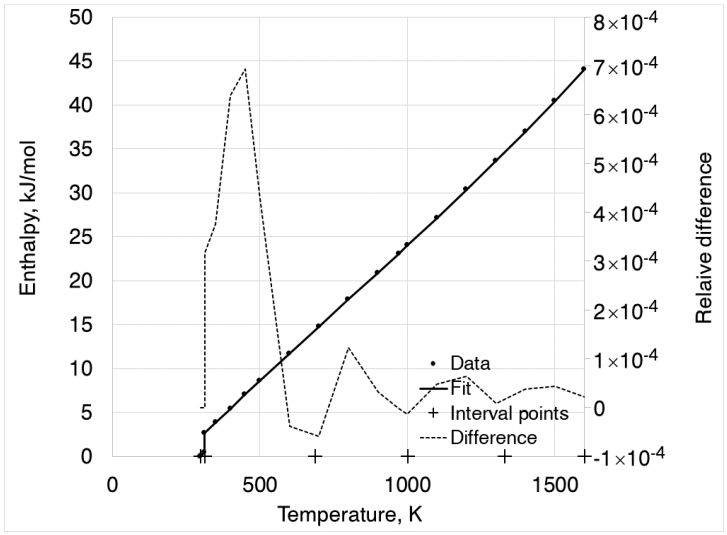


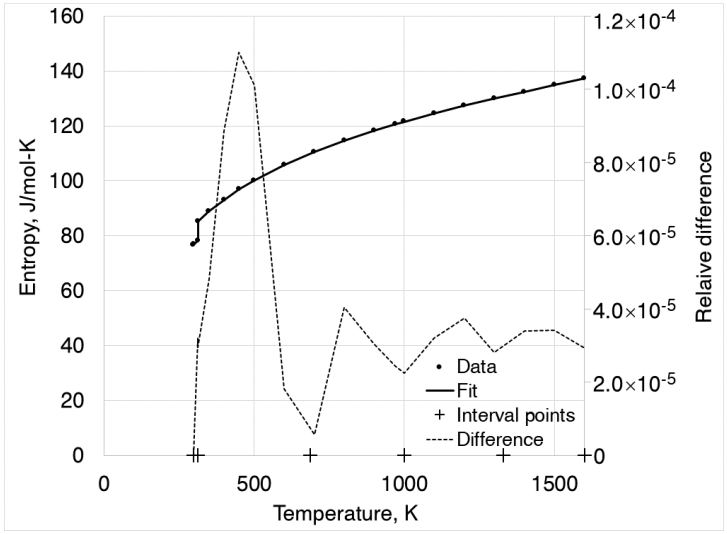


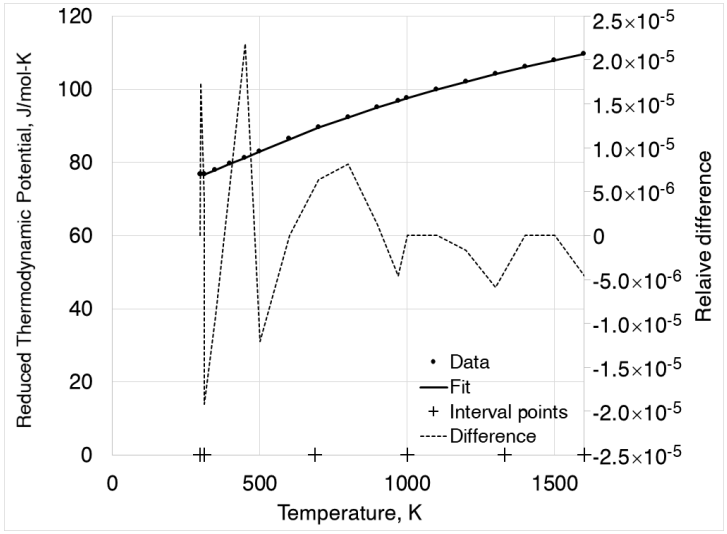
Rubidium (Rb) – Rb1(cr,l)

MW = 85.4678 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 76.778 J/mol-K – p^o = 0.1 MPa



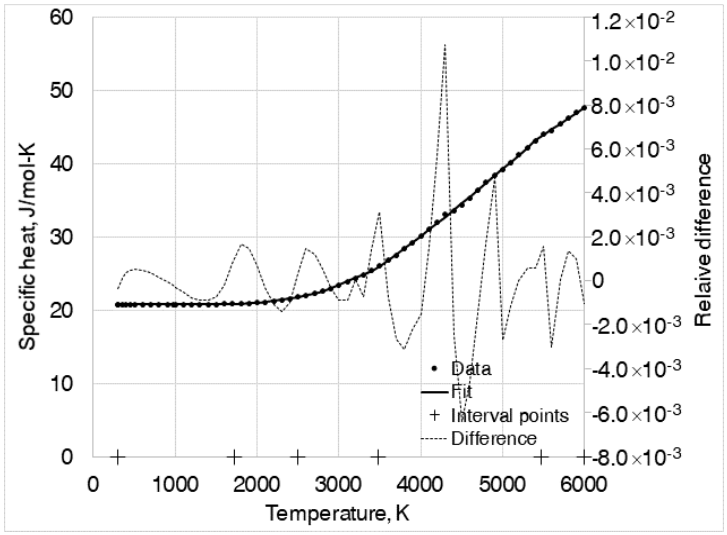


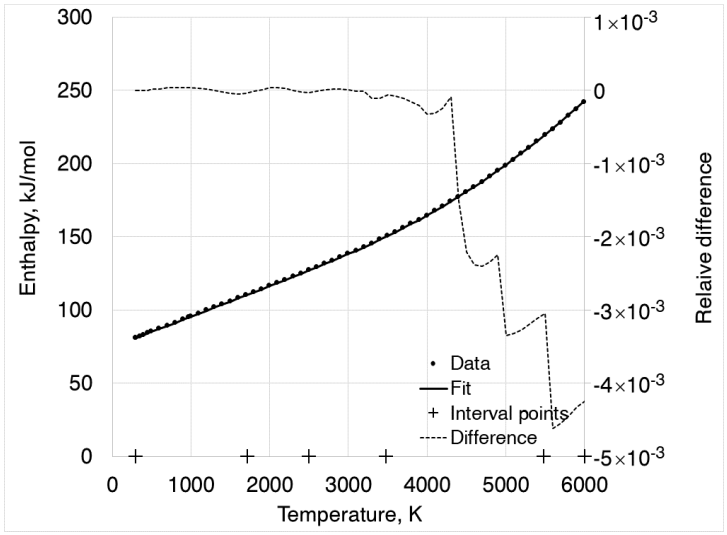


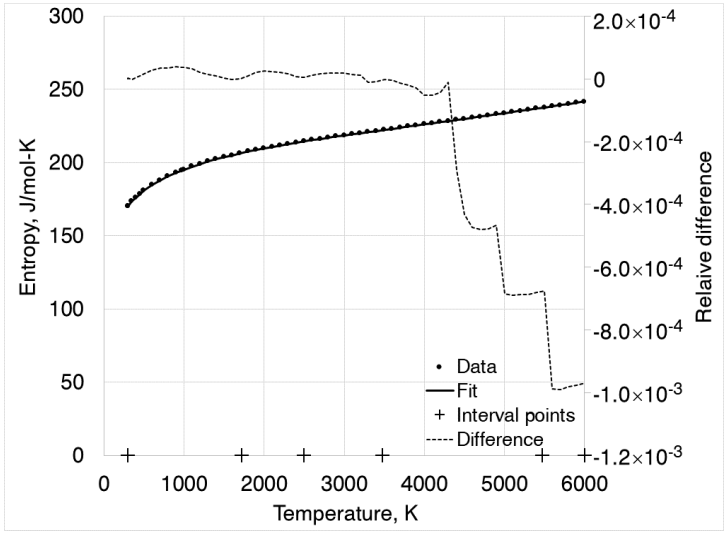


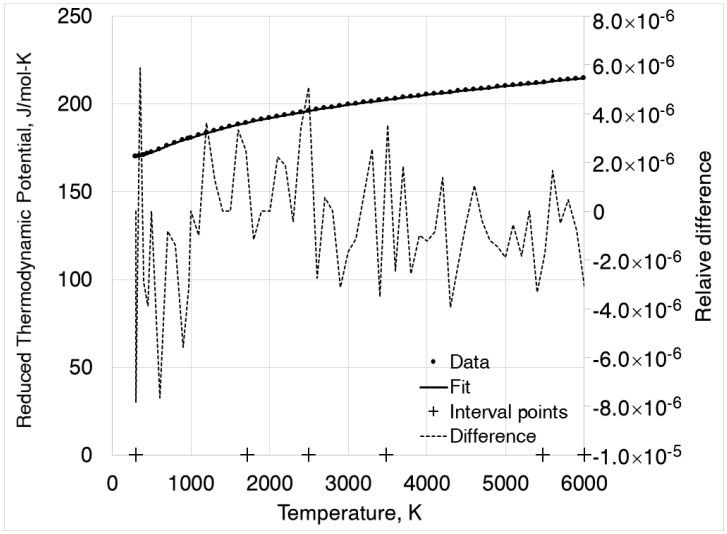
Rubidium (Rb) – Rb1(g)

MW = 85.4678 g/mol – $\Delta_f H^o$(298.15 K) = 80.9 kJ/mol – S^o(298.15 K) = 170.093 J/mol-K – p^o = 0.1 MPa



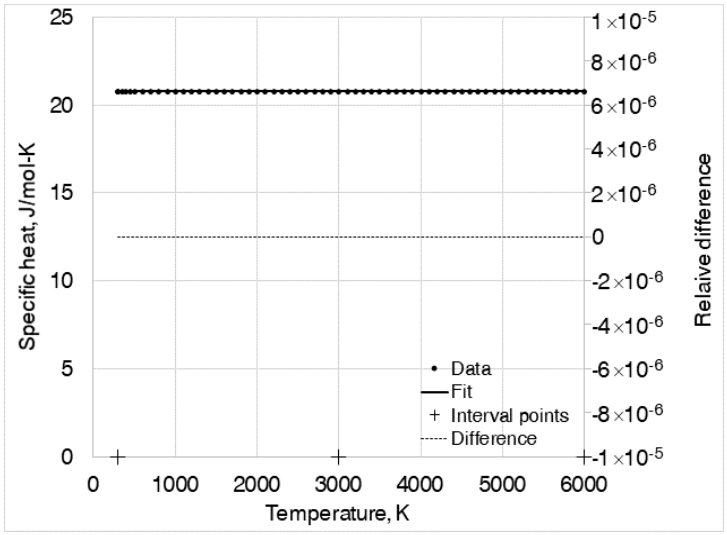


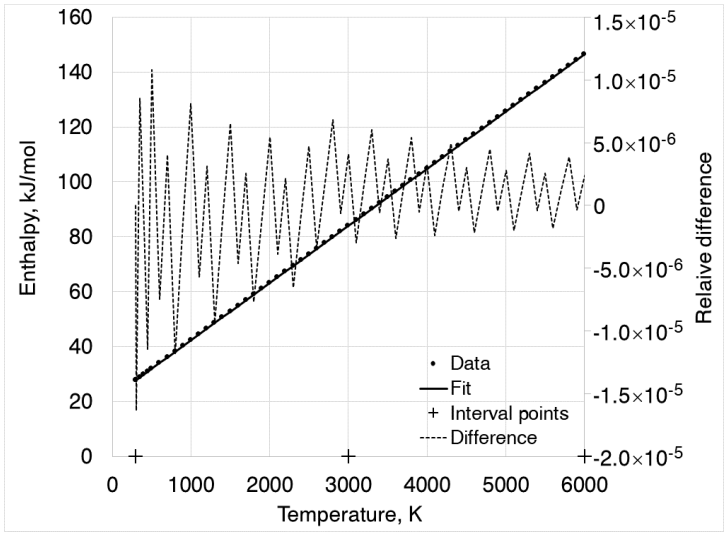


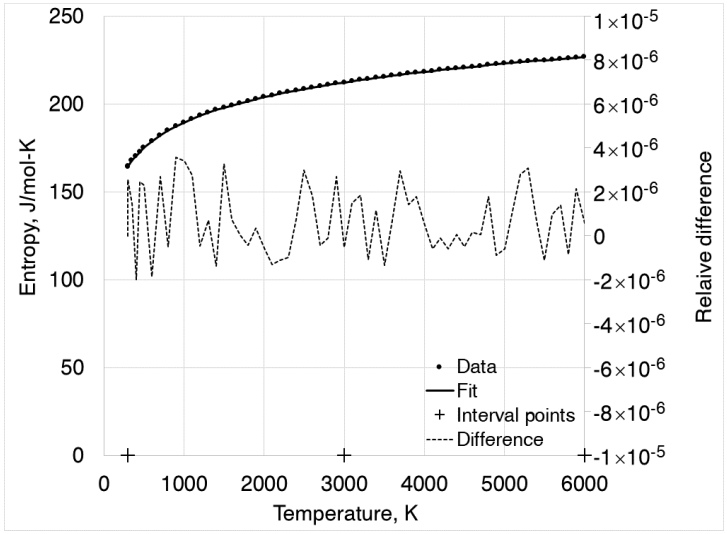


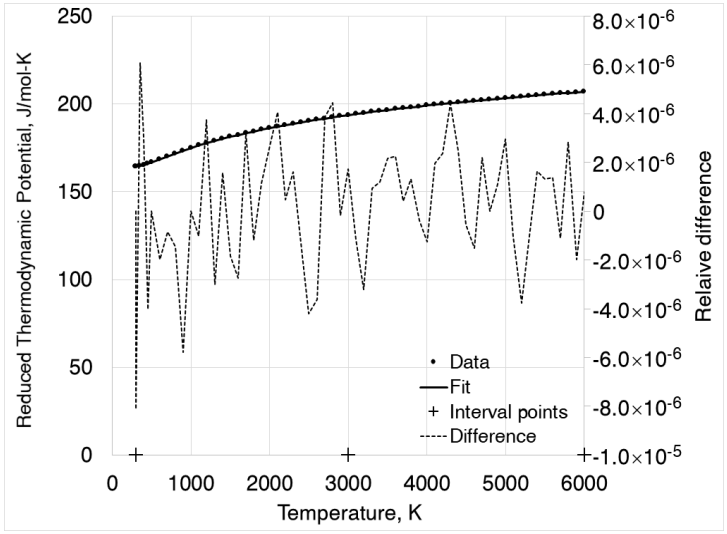
Rubidium, Ion (Rb^-) – Rb1-(g)

MW = 85.4683 g/mol – $\Delta_f H^o$(298.15 K) = 27.819 kJ/mol – S^o(298.15 K) = 164.33 J/mol-K – p^o = 0.1 MPa



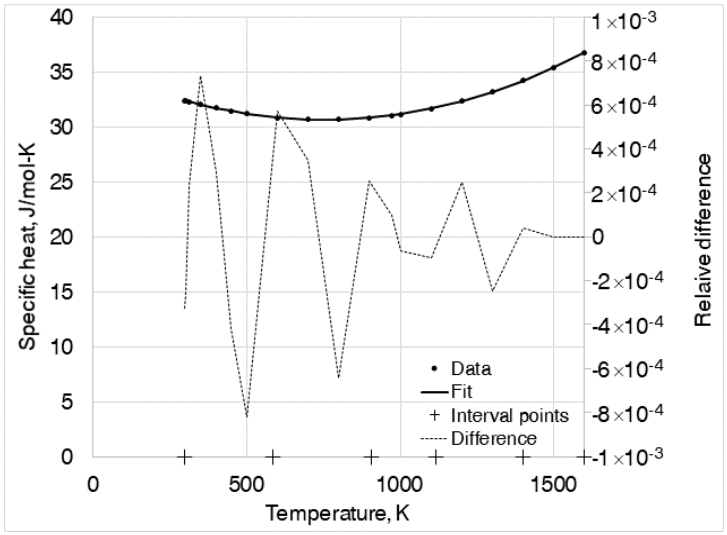


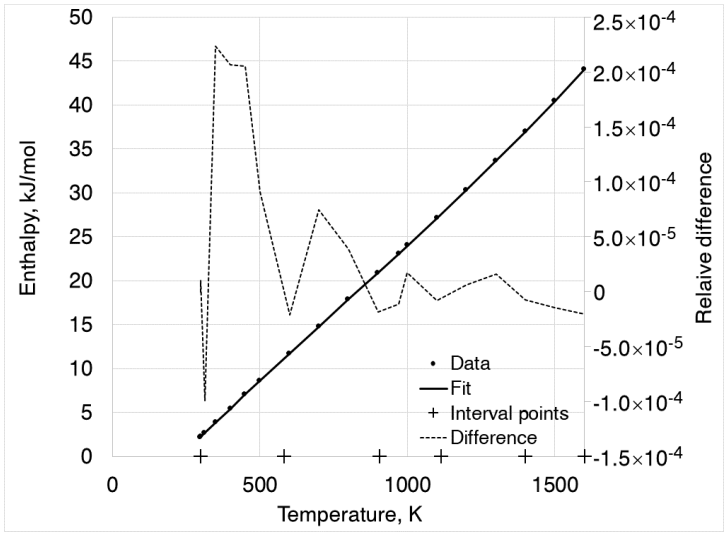


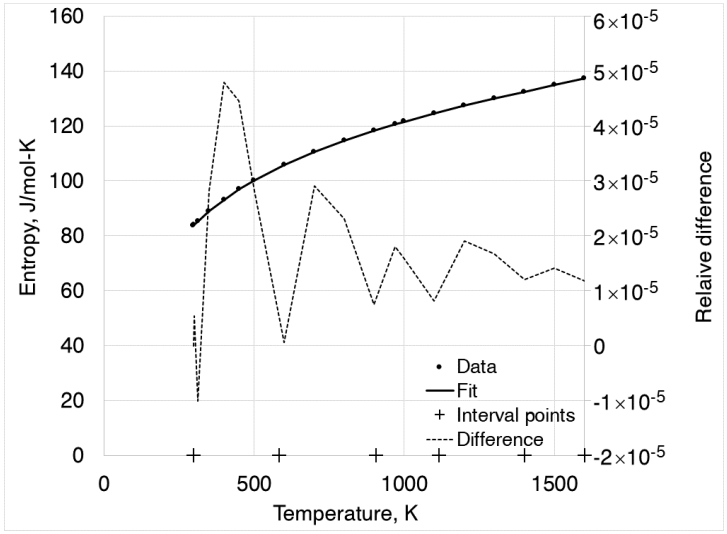


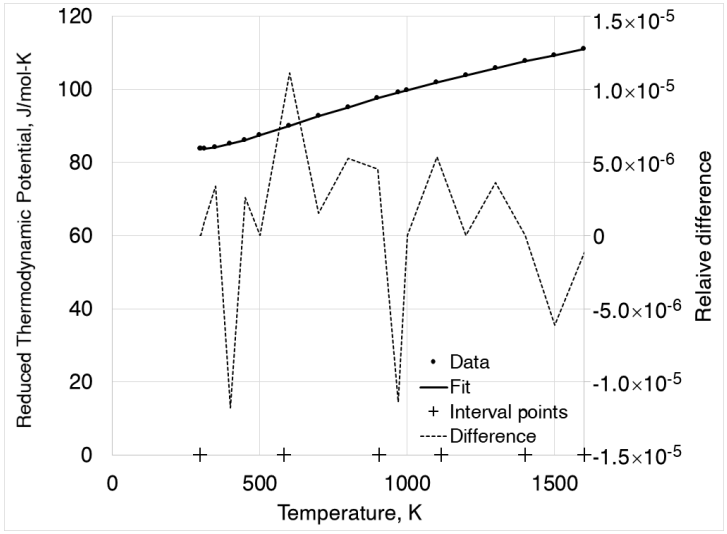
Rubidium (Rb) – Rb1(l)

MW = 85.4678 g/mol – $\Delta_f H^o$(298.15 K) = 2.184 kJ/mol – S^o(298.15 K) = 83.764 J/mol-K – p^o = 0.1 MPa



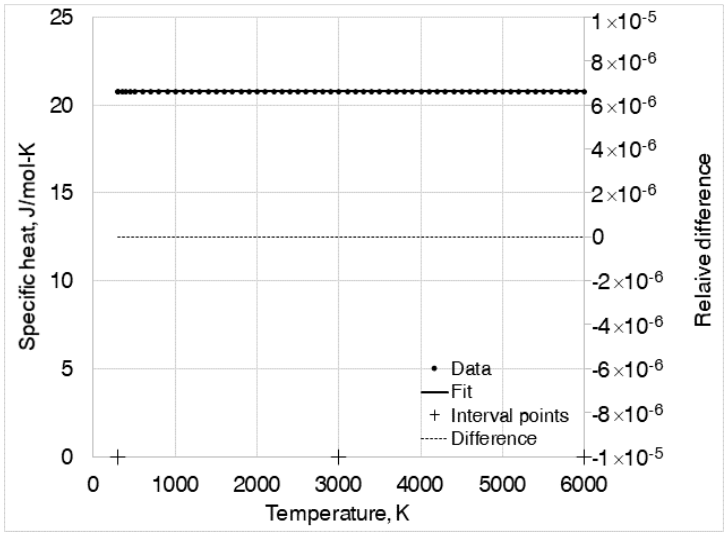


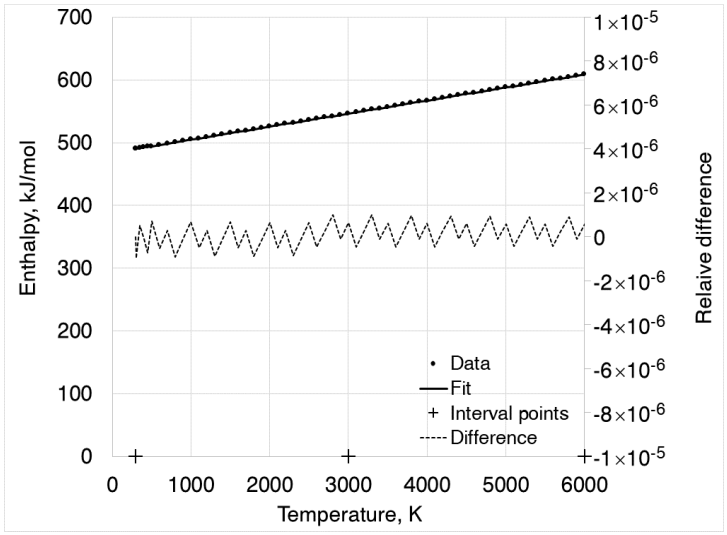


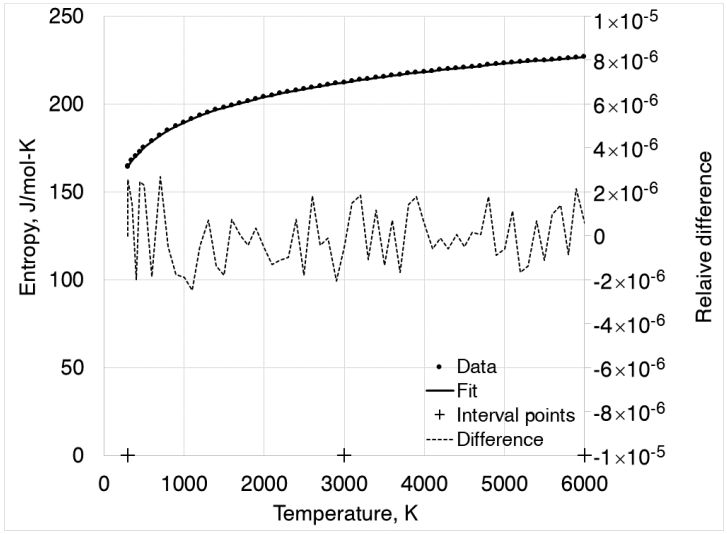


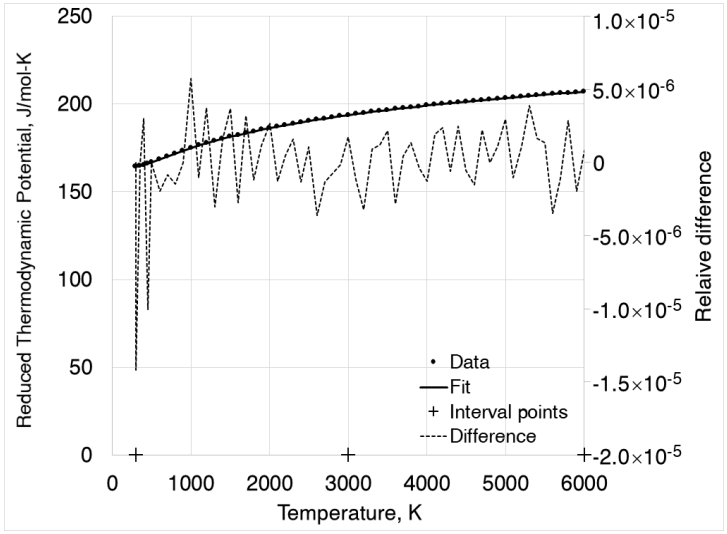
Rubidium, Ion (Rb^+) – Rb1+(g)

MW = 85.4673 g/mol – $\Delta_f H^o$(298.15 K) = 490.129 kJ/mol – S^o(298.15 K) = 164.33 J/mol-K – p^o = 0.1 MPa



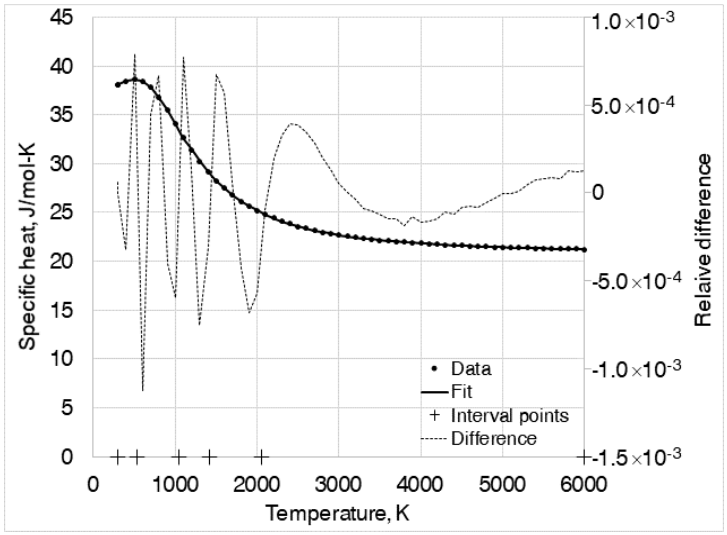


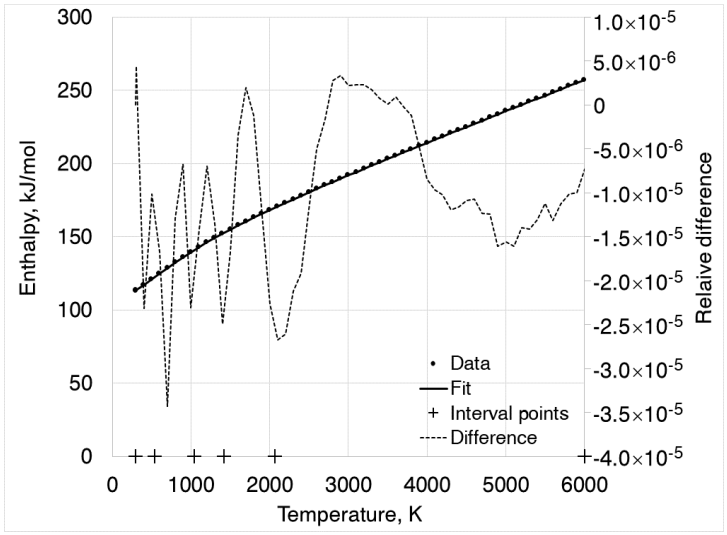


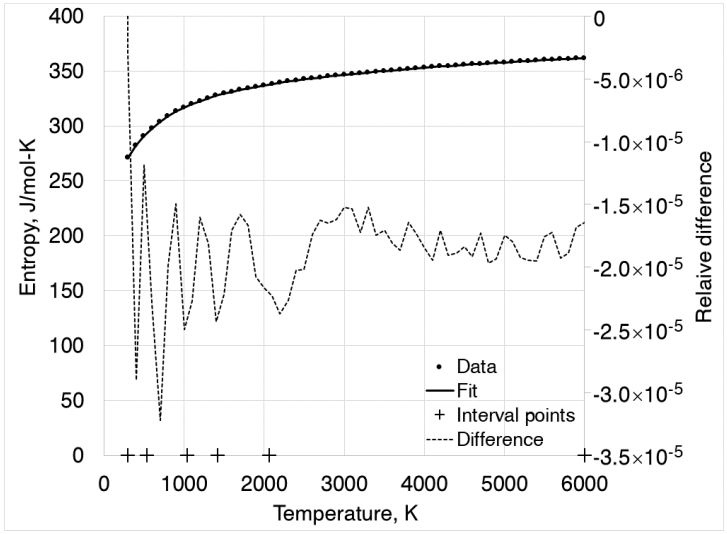


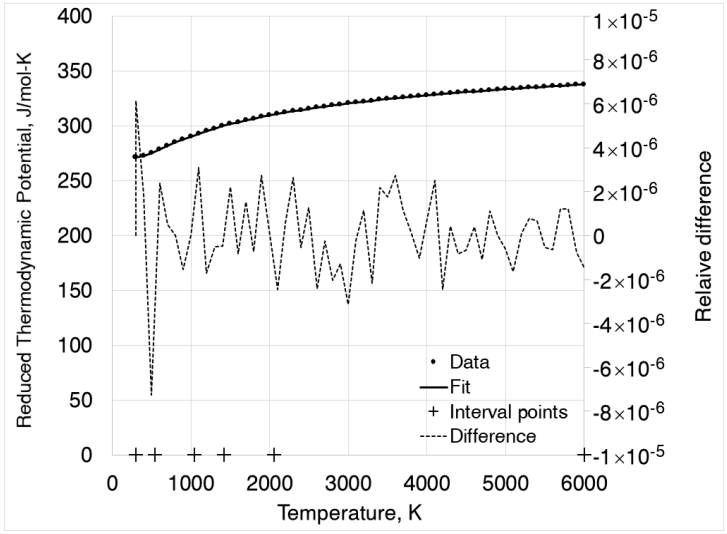
Rubidium (Rb_2) – Rb2(g)

MW = 170.9356 g/mol – $\Delta_f H^o$(298.15 K) = 113.29 kJ/mol – S^o(298.15 K) = 271.067 J/mol-K – p^o = 0.1 MPa



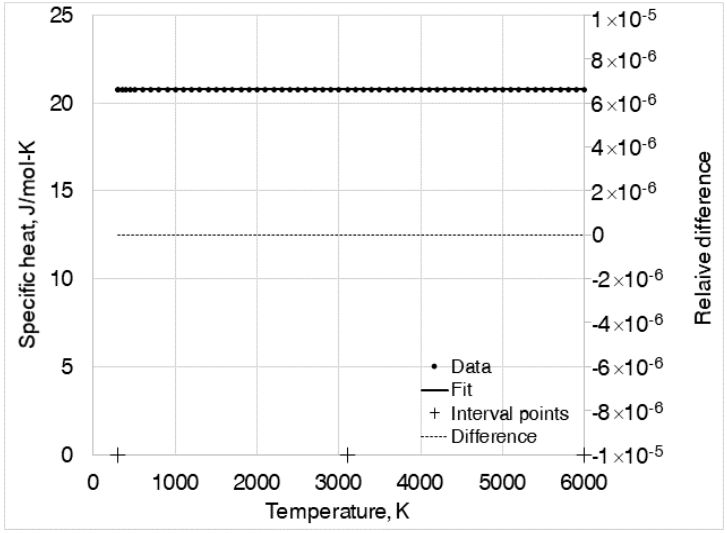


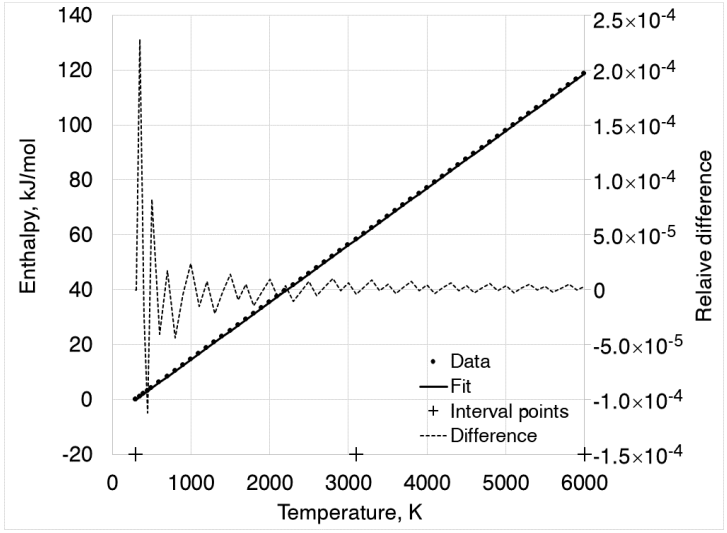


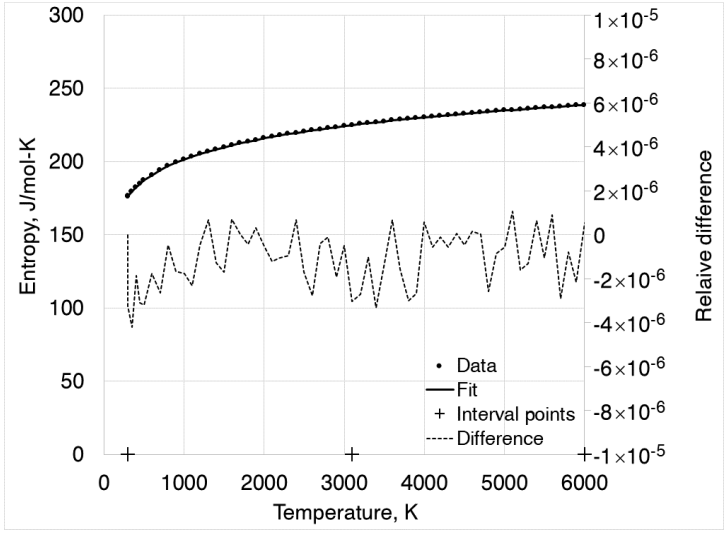


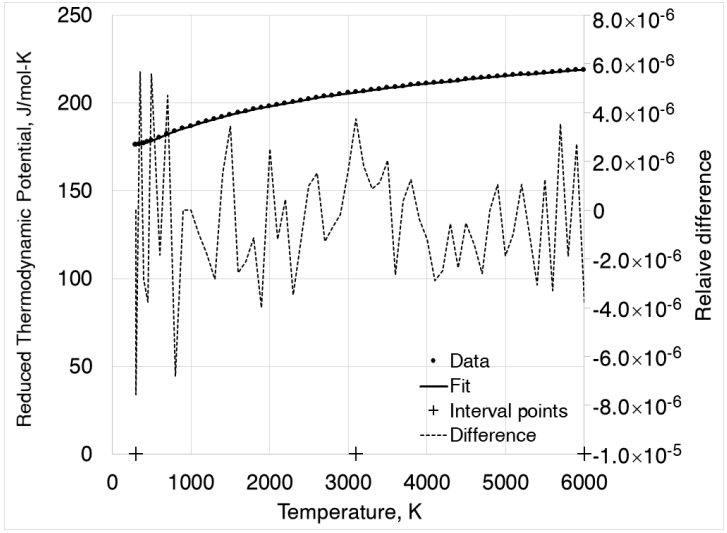
Radon (Rn) – Rn1(ref)

MW = 222 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 176.235 J/mol-K – p^o = 0.1 MPa



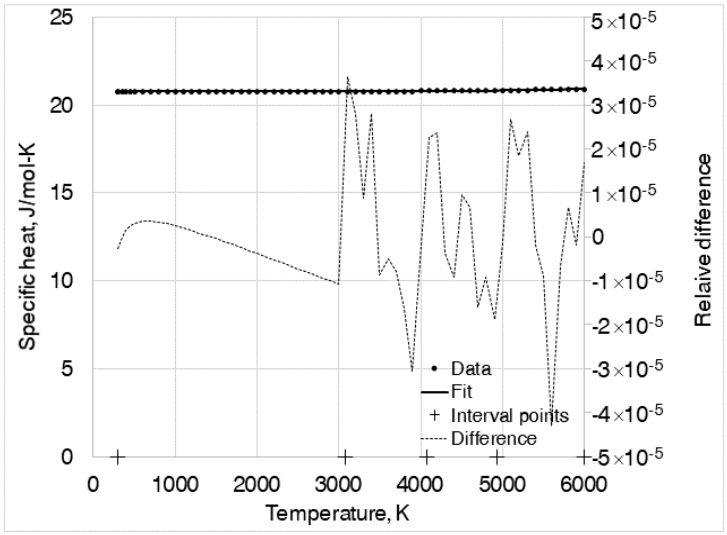


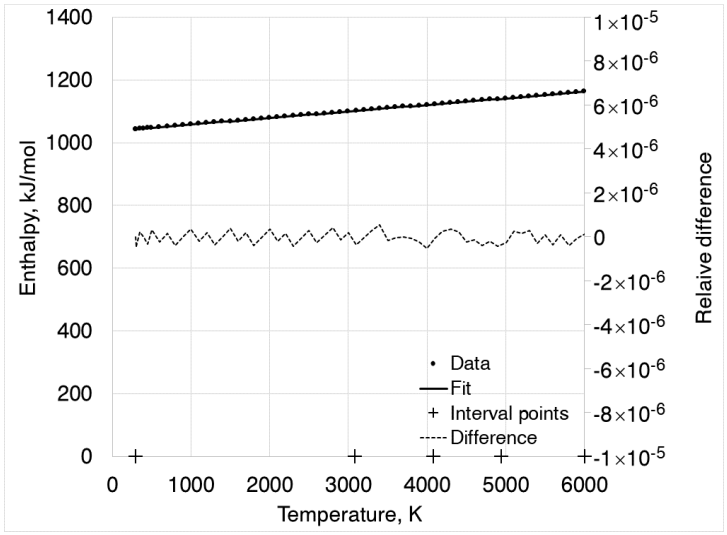


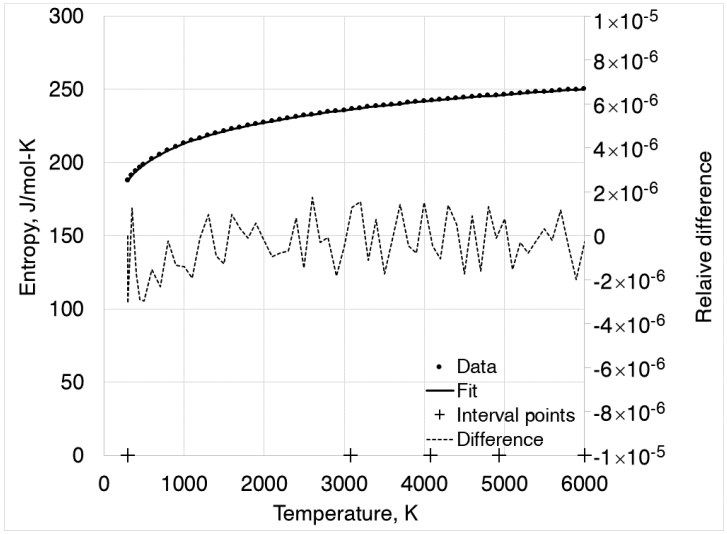


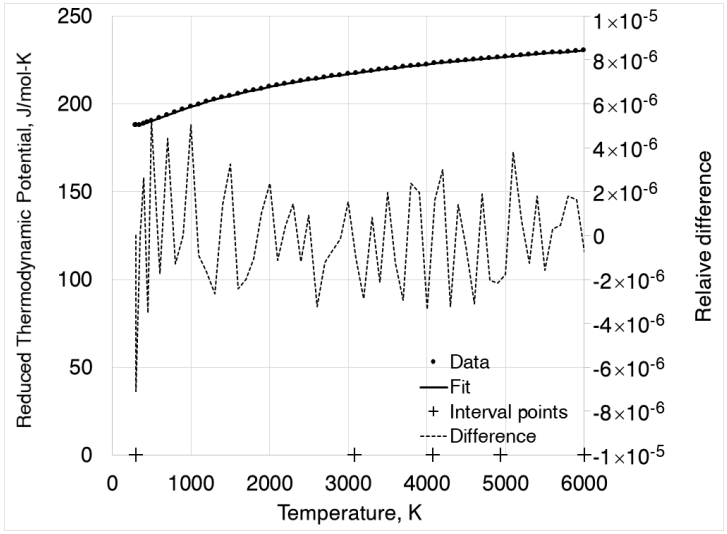
Radon, Ion (Rn^+) – Rn1+(g)

MW = 222 g/mol – $\Delta_f H^o$(298.15 K) = 1043.27 kJ/mol – S^o(298.15 K) = 187.761 J/mol-K – p^o = 0.1 MPa



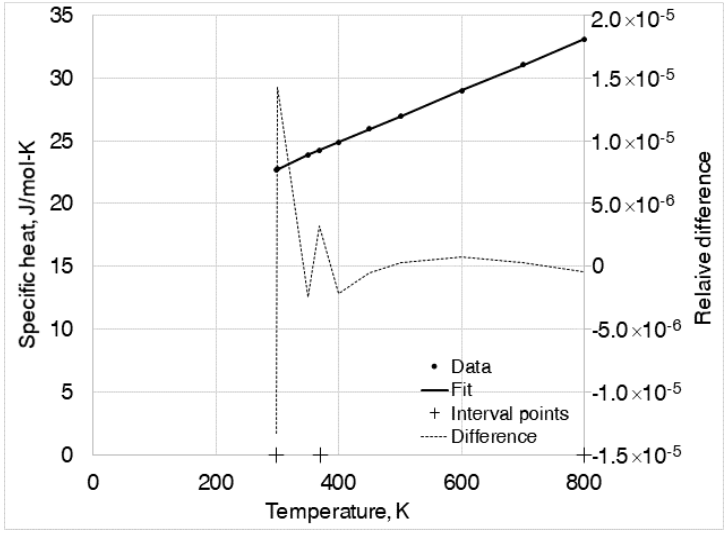


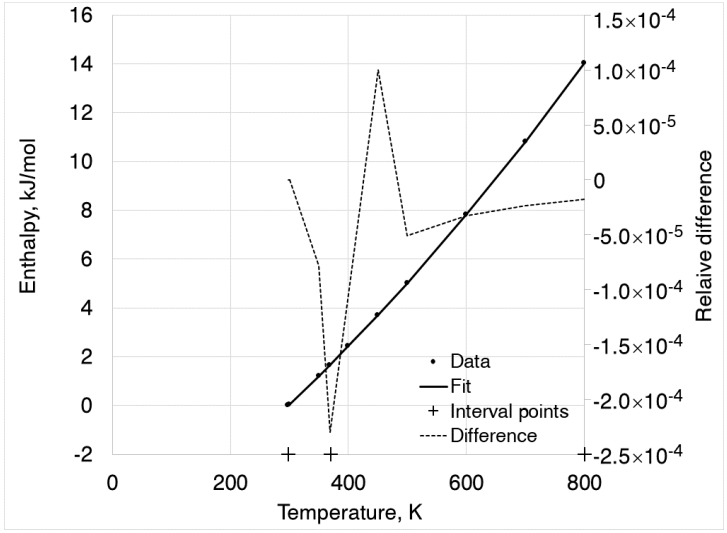


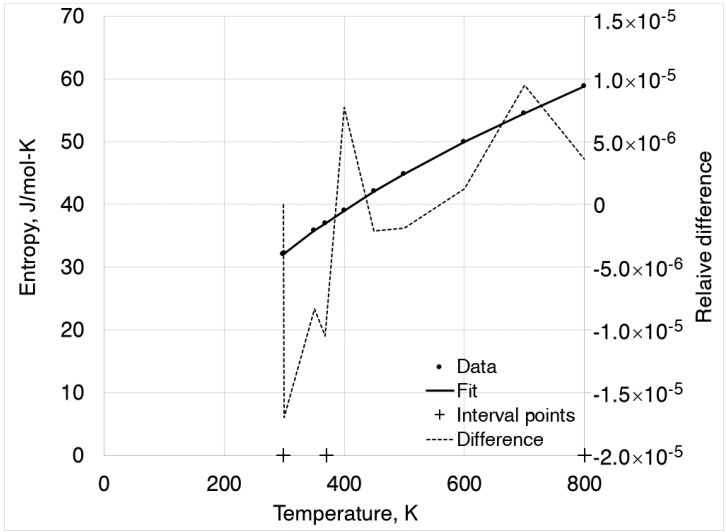


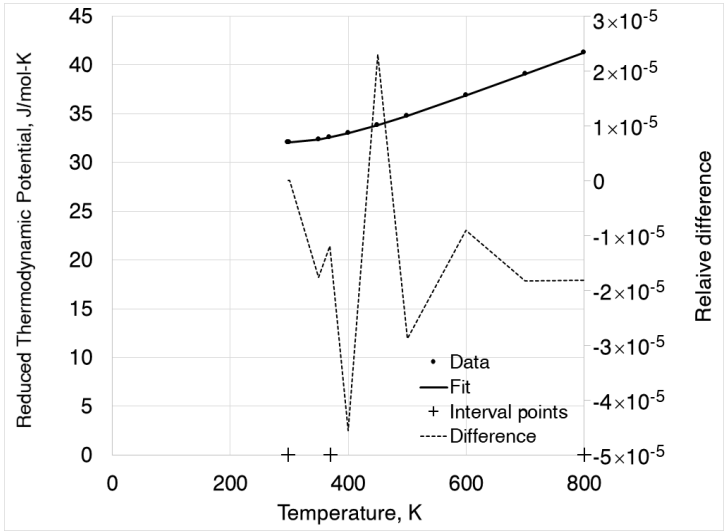
Sulfur, Orthorhombic (S) – S1(cr)_01

MW = 32.06 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 32.056 J/mol-K – p^o = 0.1 MPa



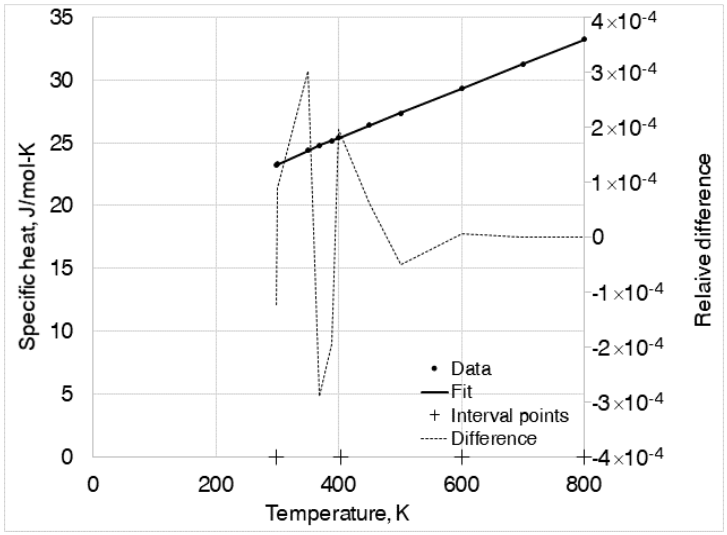


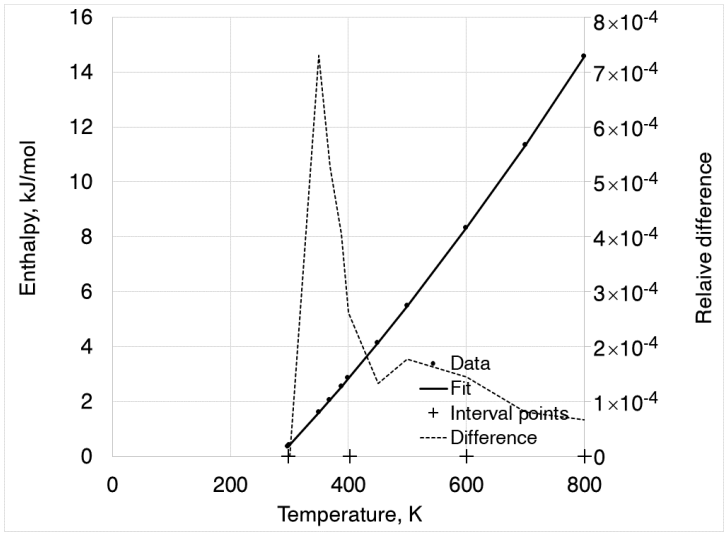


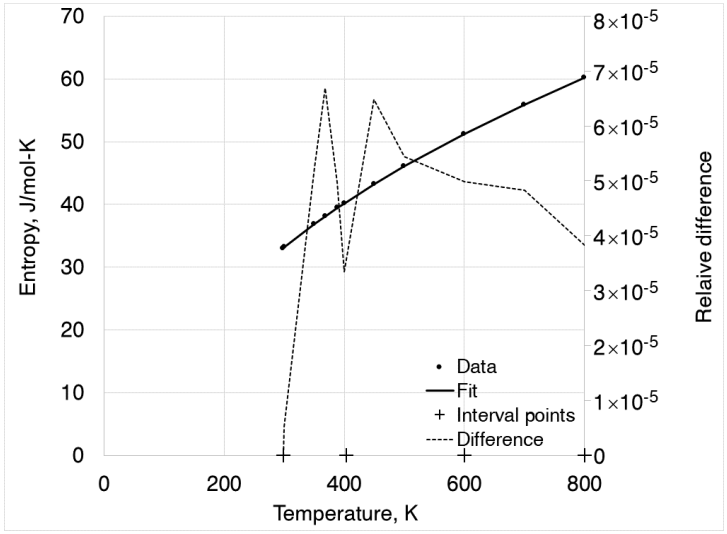


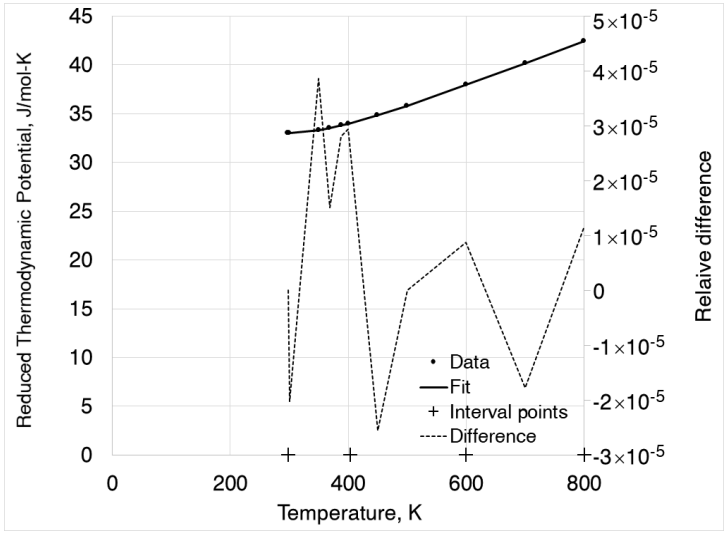
Sulfur, Monoclinic (S) – S1(cr)_02

MW = 32.06 g/mol – $\Delta_f H^o$(298.15 K) = .36 kJ/mol – S^o(298.15 K) = 33.028 J/mol-K – p^o = 0.1 MPa



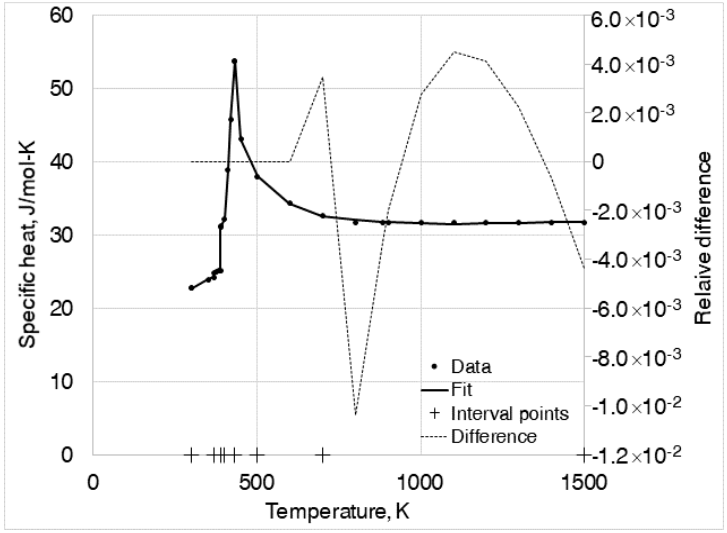


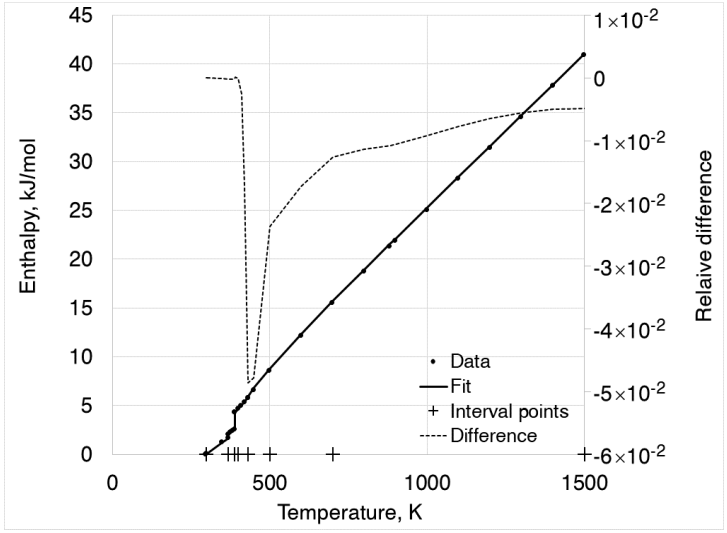


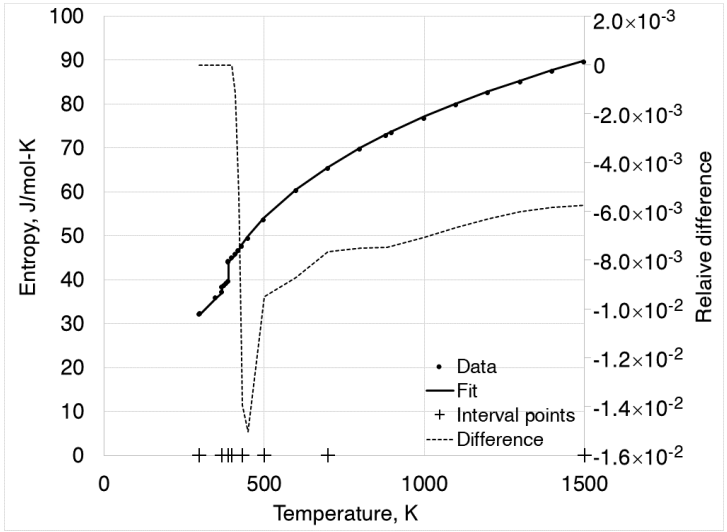


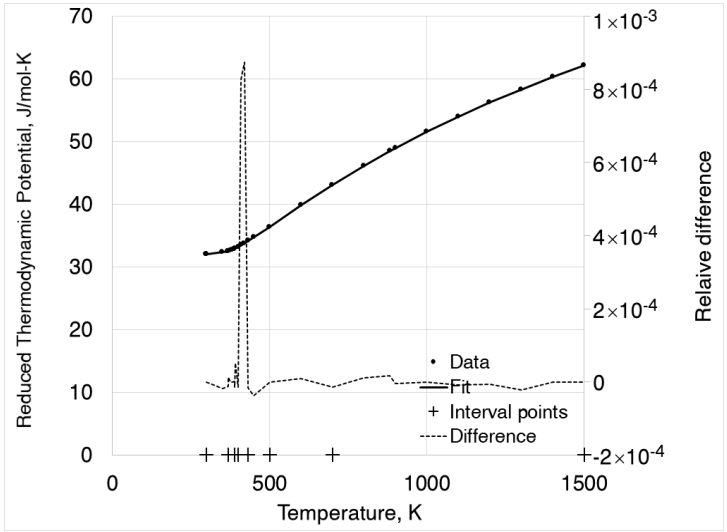
Sulfur (S) – S1(cr,l)

MW = 32.06 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 32.056 J/mol-K – p^o = 0.1 MPa



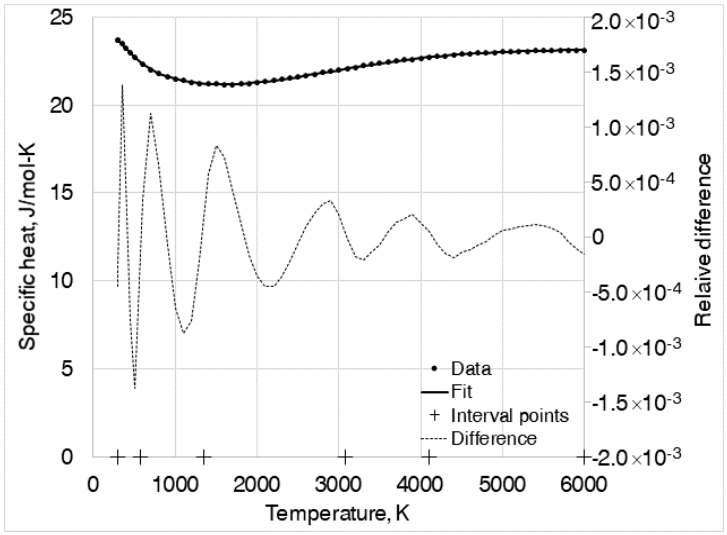


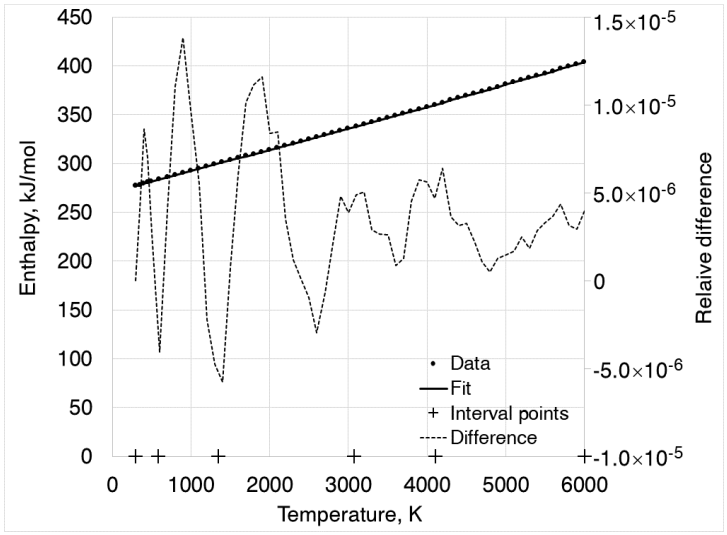


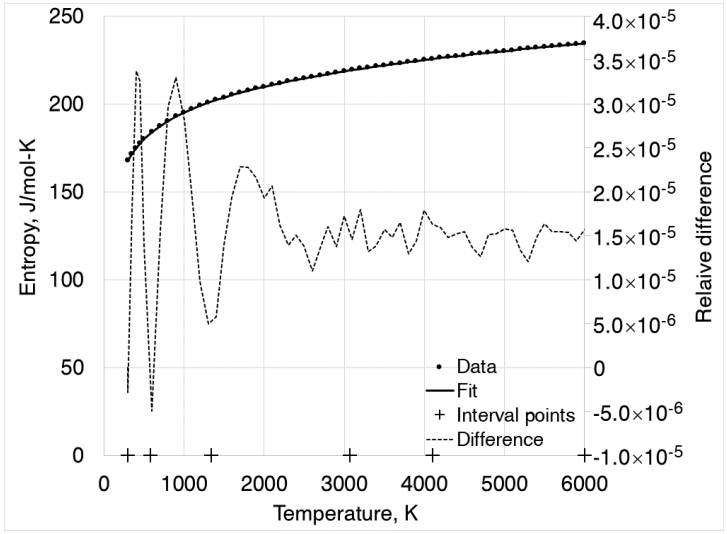


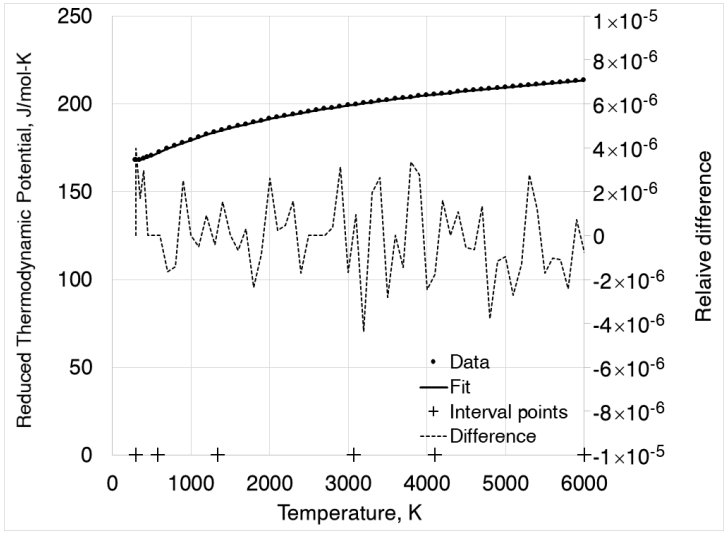
Sulfur (S) – S1(g)

MW = 32.06 g/mol – $\Delta_f H^o$(298.15 K) = 276.98 kJ/mol – S^o(298.15 K) = 167.828 J/mol-K – p^o = 0.1 MPa



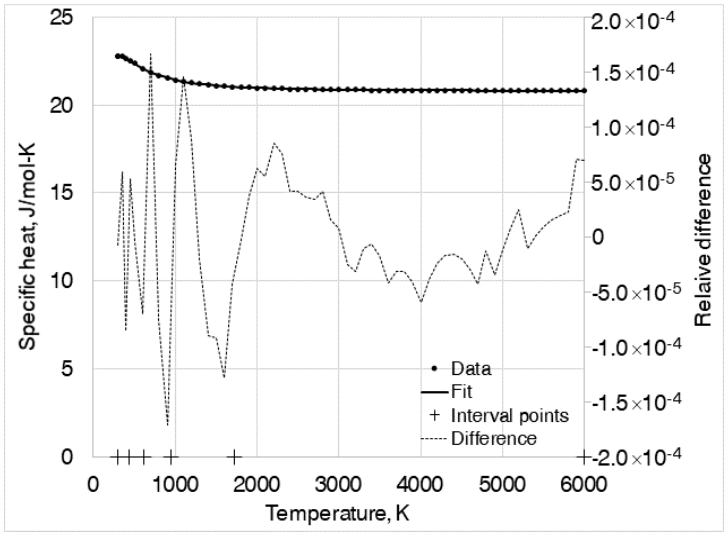


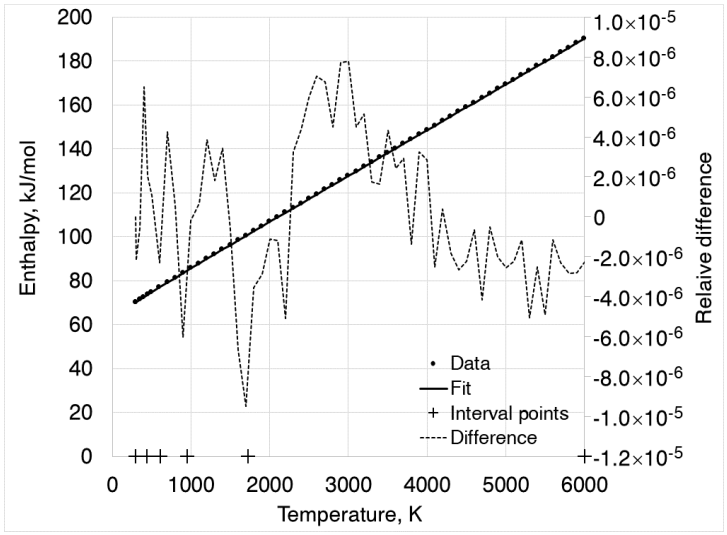


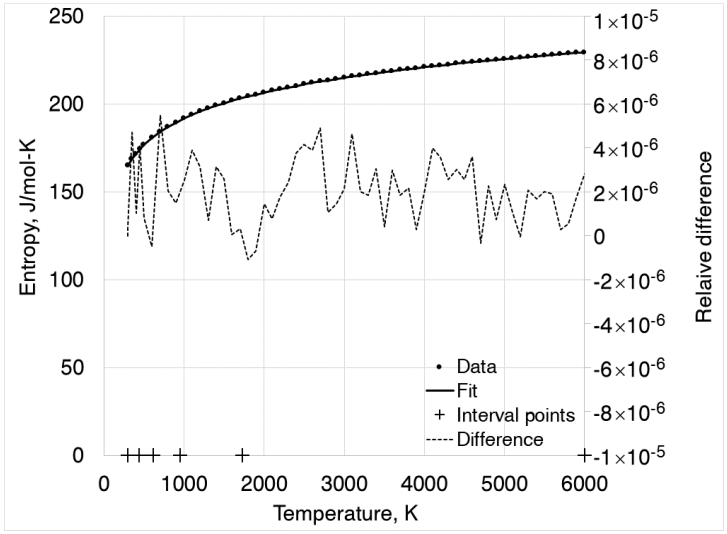


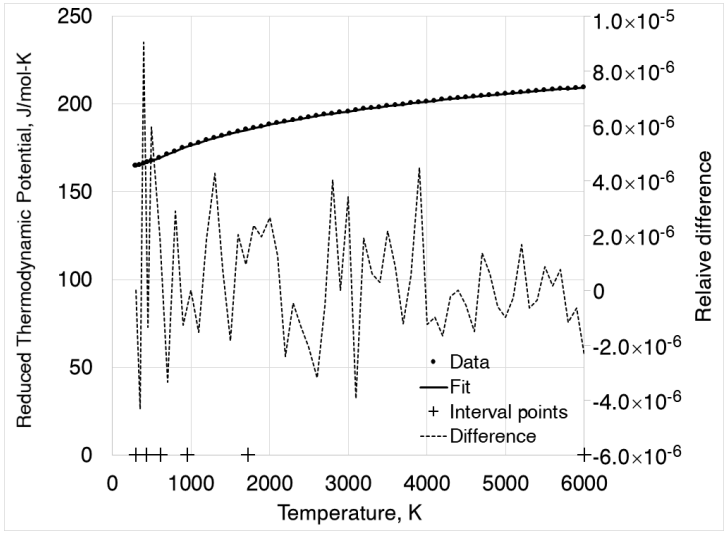
Sulfur, Ion (S^-) – S1-(g)

MW = 32.06 g/mol – $\Delta_f H^o$(298.15 K) = 70.18 kJ/mol – S^o(298.15 K) = 164.918 J/mol-K – p^o = 0.1 MPa



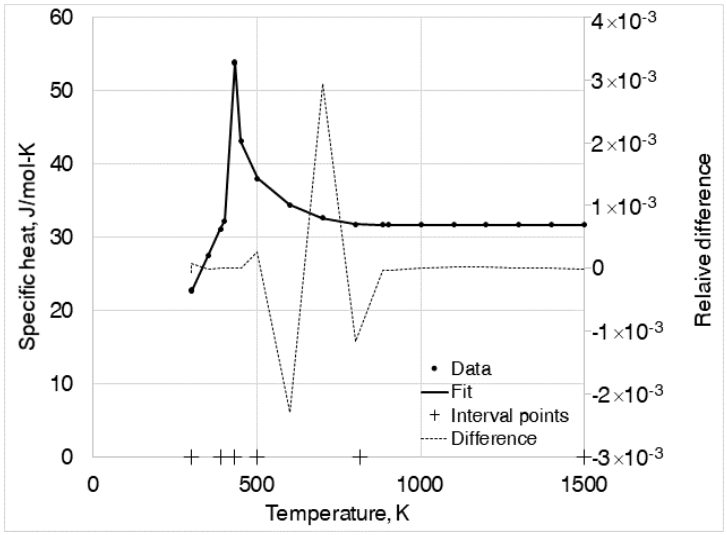


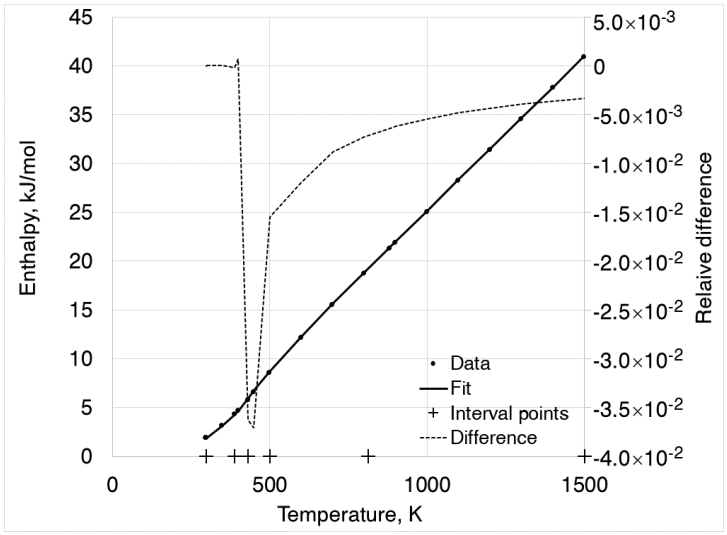


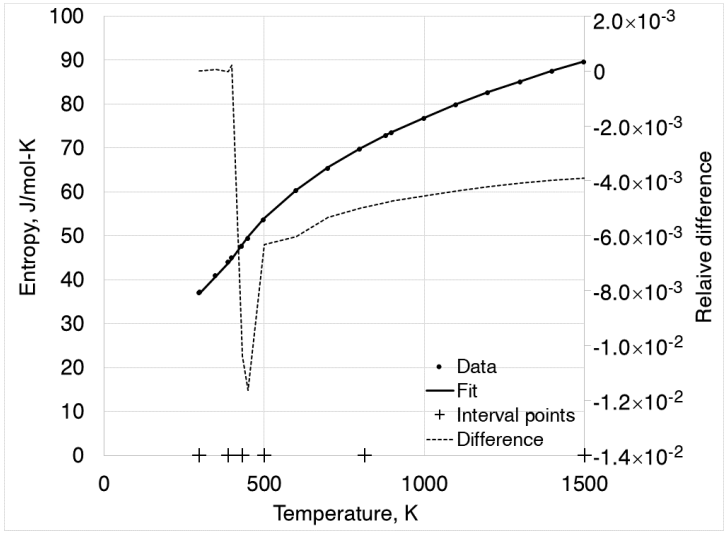


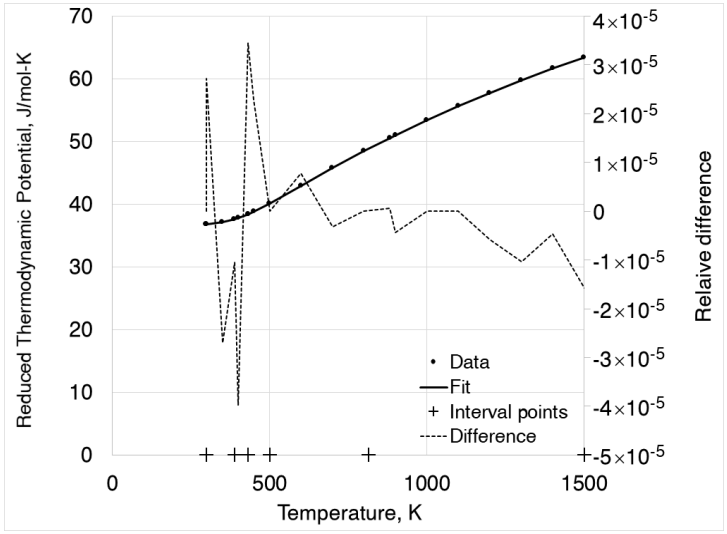
Sulfur (S) – S1(l)

MW = 32.06 g/mol – $\Delta_f H^o$(298.15 K) = 1.854 kJ/mol – S^o(298.15 K) = 36.825 J/mol-K – p^o = 0.1 MPa



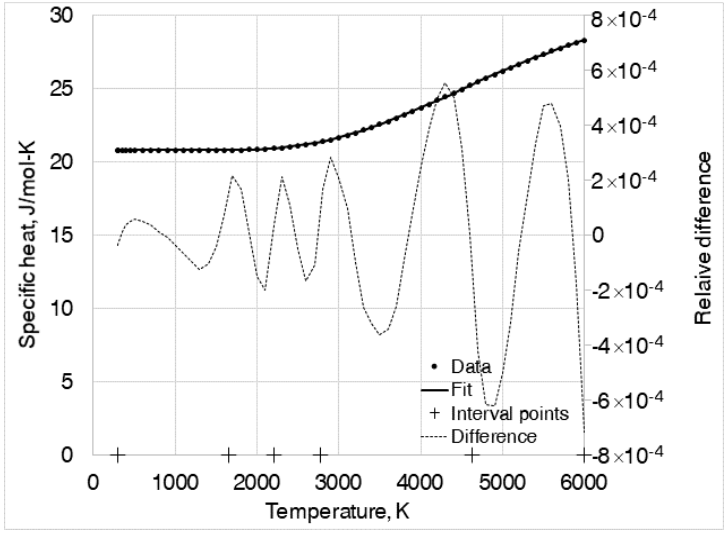


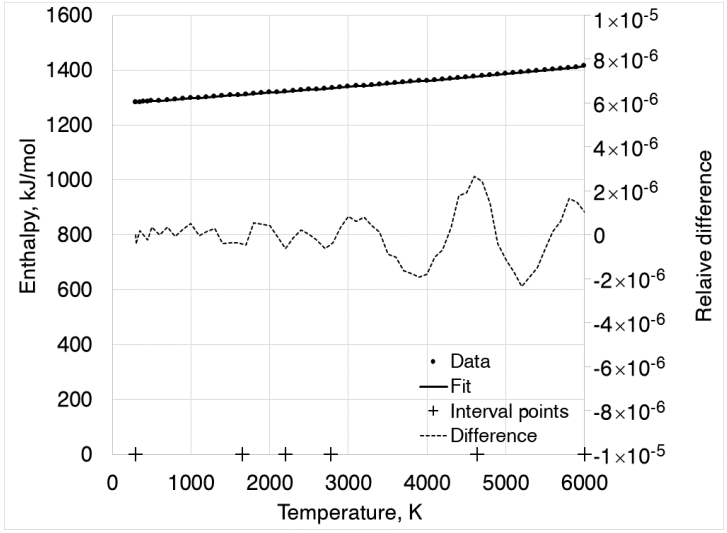


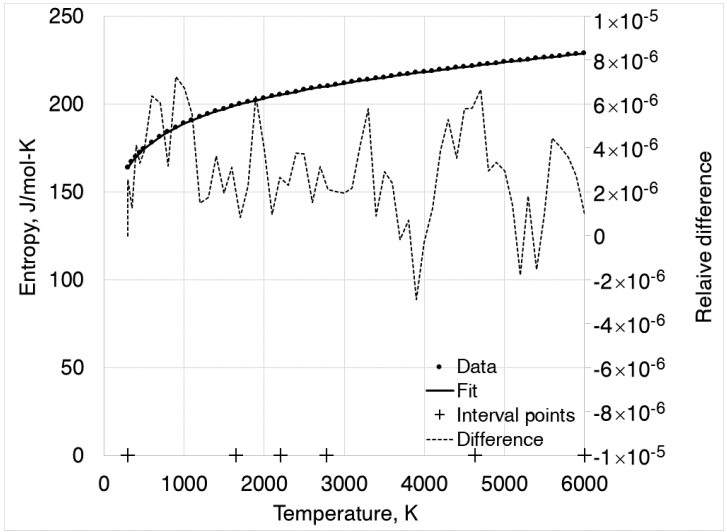


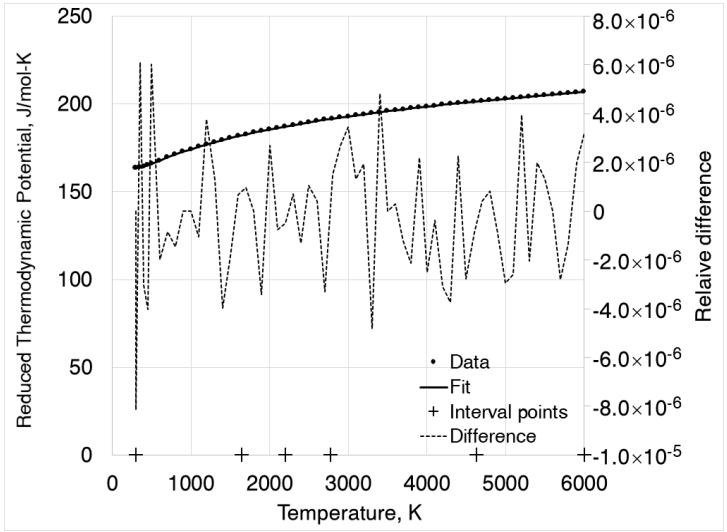
Sulfur, Ion (S^+) – S1+(g)

MW = 32.06 g/mol – $\Delta_f H^o$(298.15 K) = 1282.293 kJ/mol – S^o(298.15 K) = 163.627 J/mol-K – p^o = 0.1 MPa



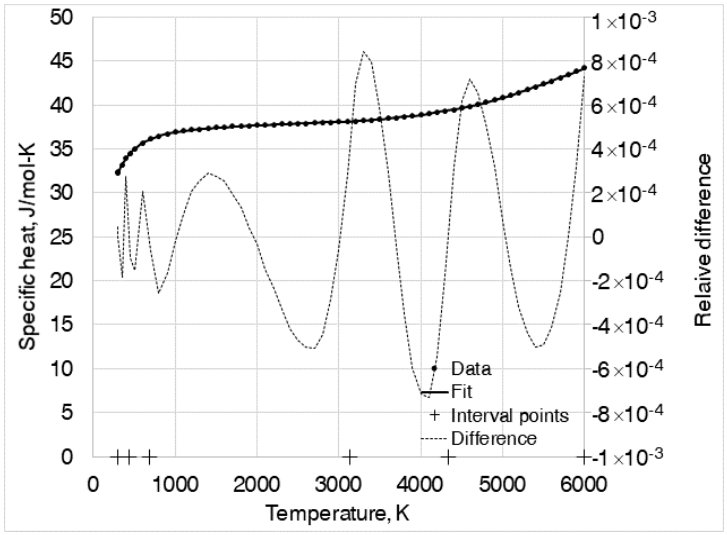


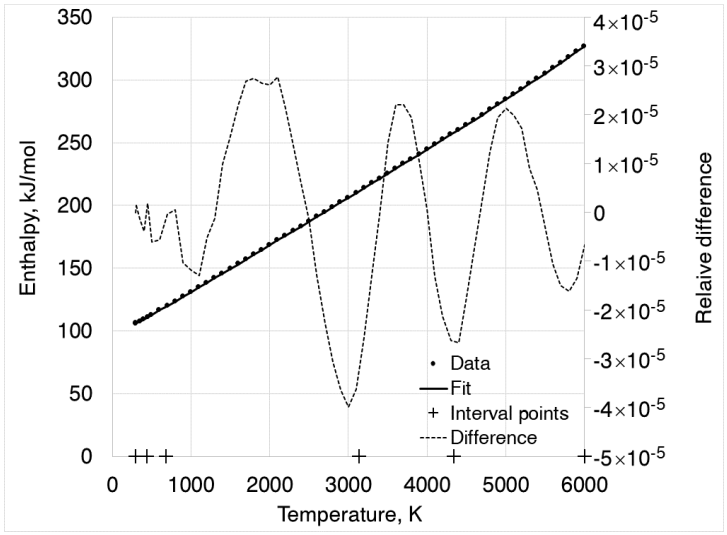


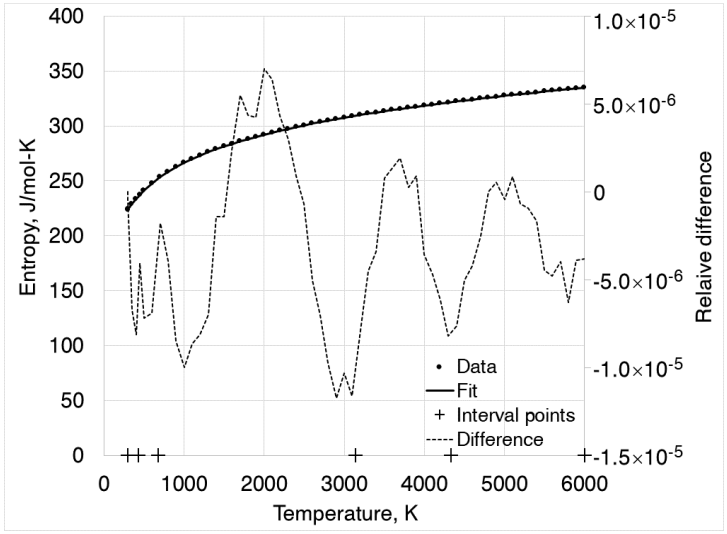


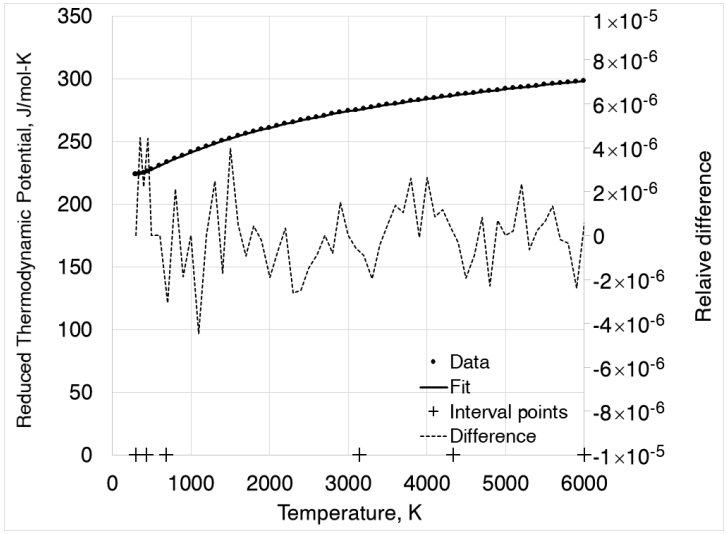
Silicon Sulfide (SiS) – S1Si1(g)

MW = 60.15 g/mol – $\Delta_f H^o$(298.15 K) = 105.96 kJ/mol – S^o(298.15 K) = 223.799 J/mol-K – p^o = 0.1 MPa



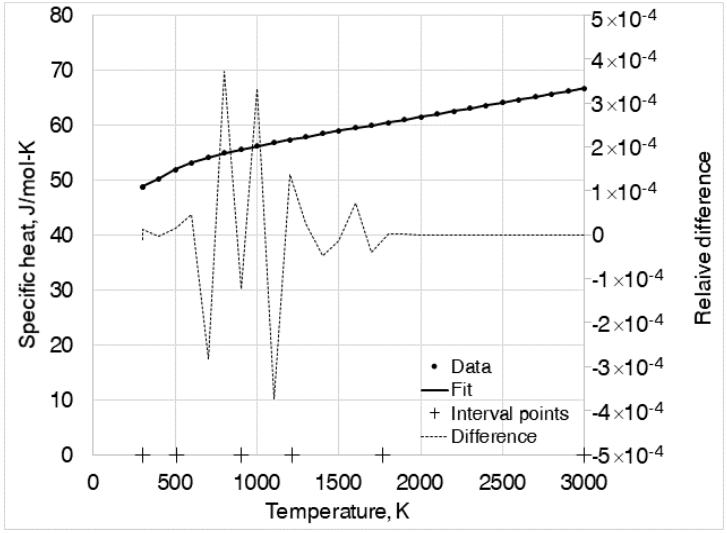


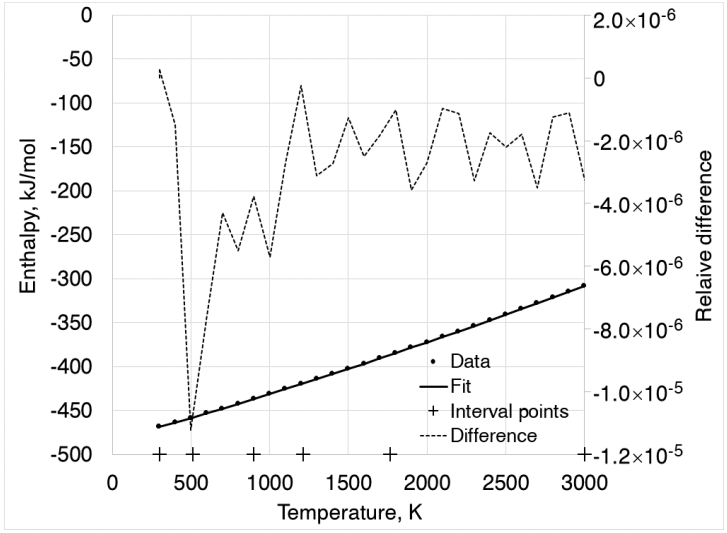


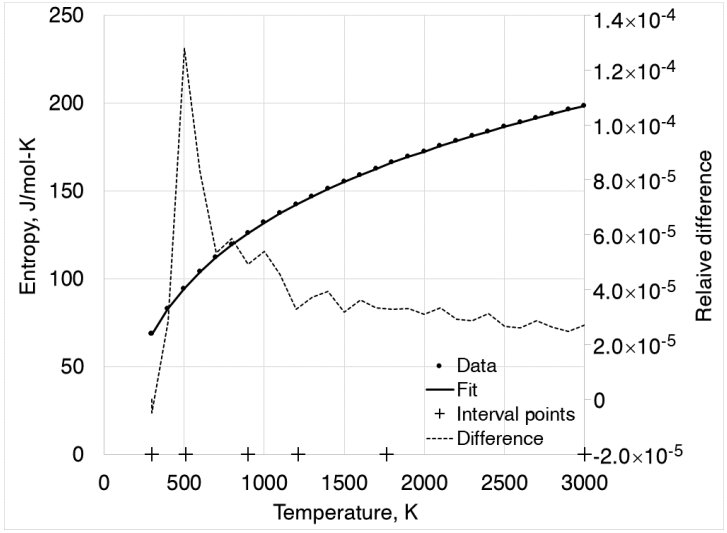


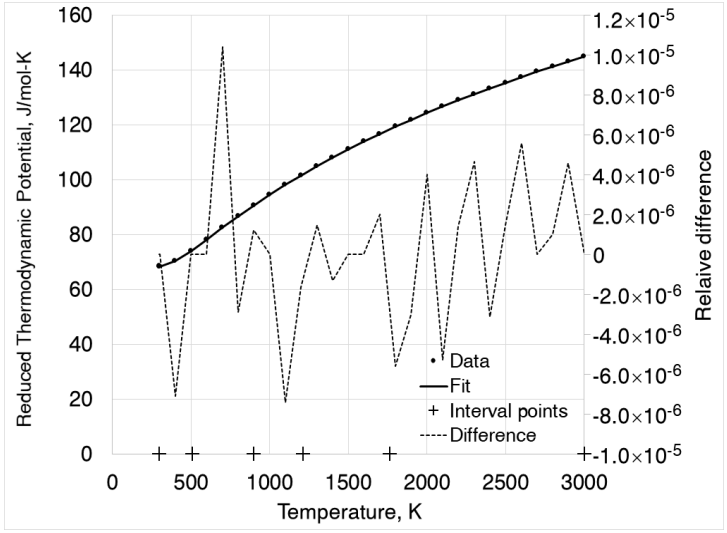
Strontium Sulfide (SrS) – S1Sr1(cr)

MW = 119.68 g/mol – $\Delta_f H^o$(298.15 K) = -468.608 kJ/mol – S^o(298.15 K) = 68.41 J/mol-K – p^o = 0.1 MPa



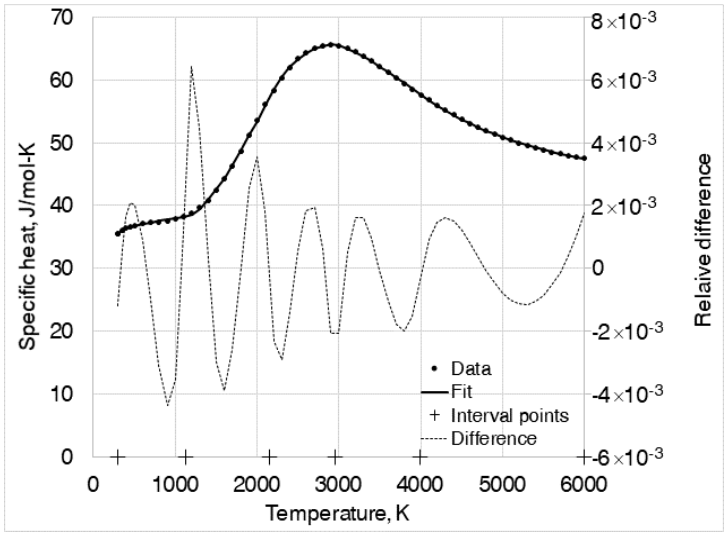


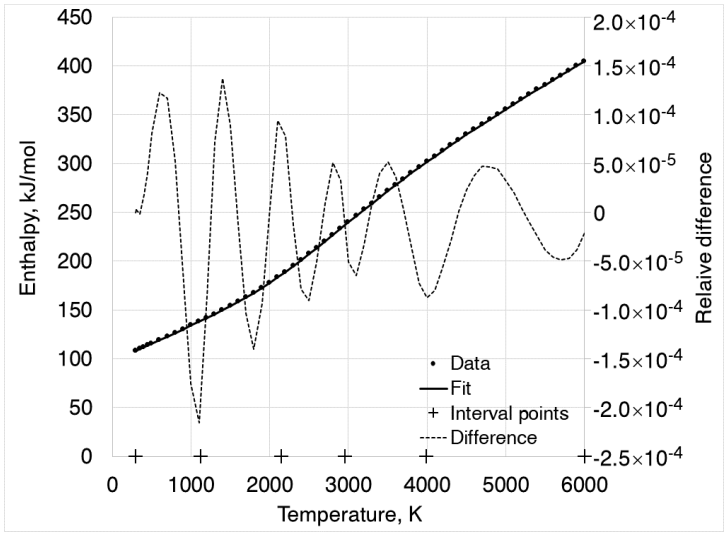


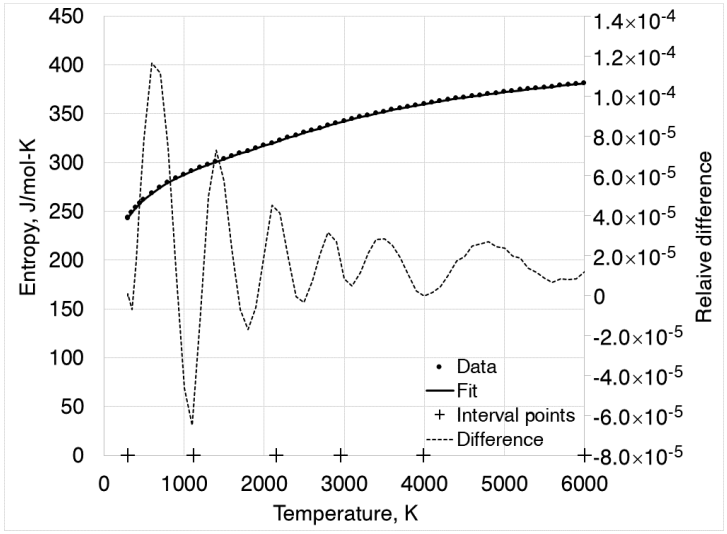


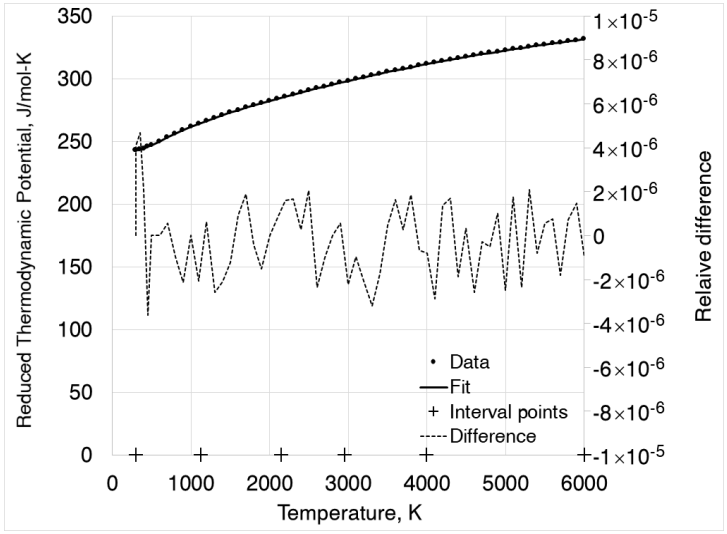
Strontium Sulfide (SrS) – S1Sr1(g)

MW = 119.68 g/mol – $\Delta_f H^o$(298.15 K) = 108.198 kJ/mol – S^o(298.15 K) = 243.107 J/mol-K – p^o = 0.1 MPa



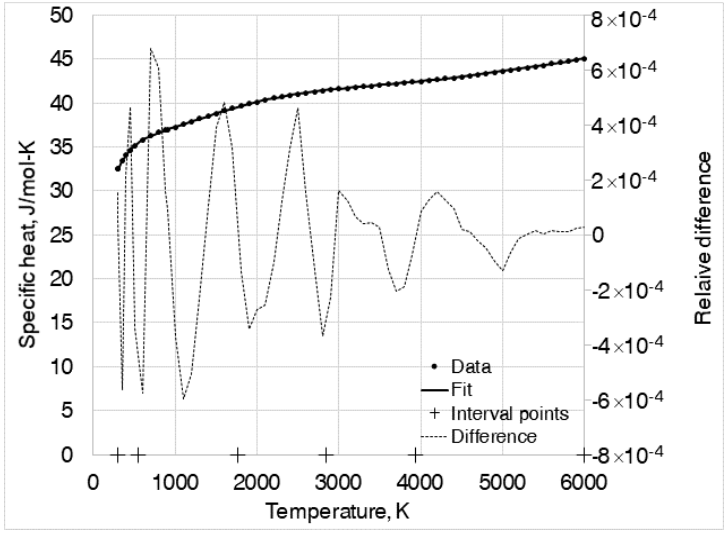


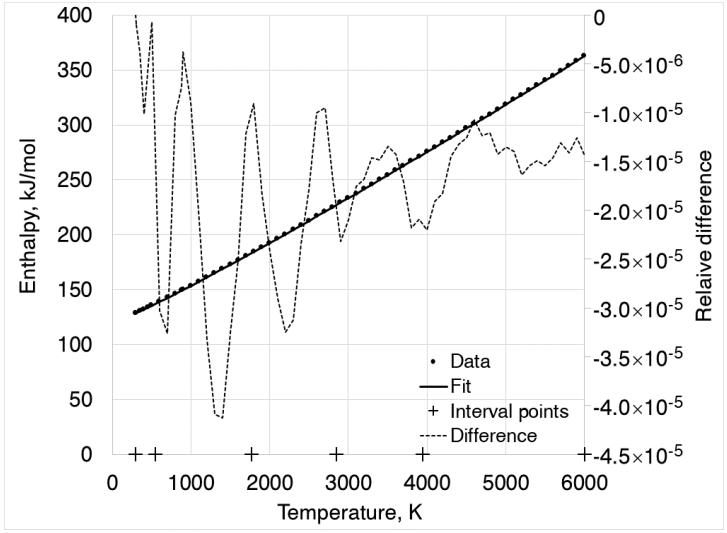


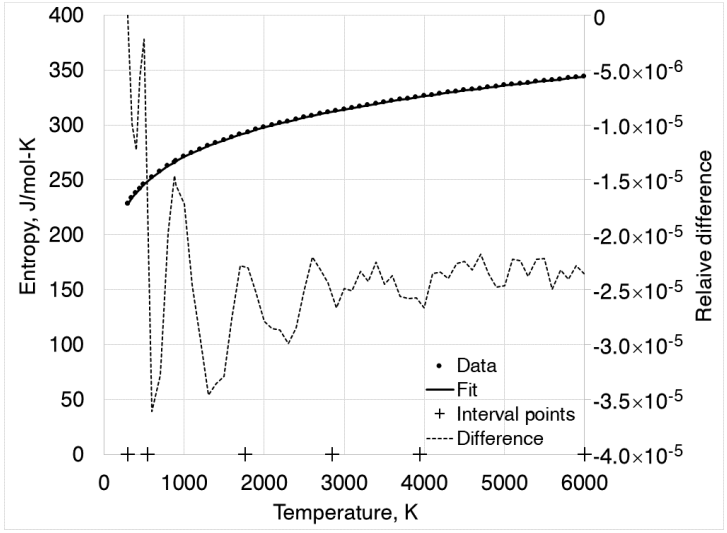


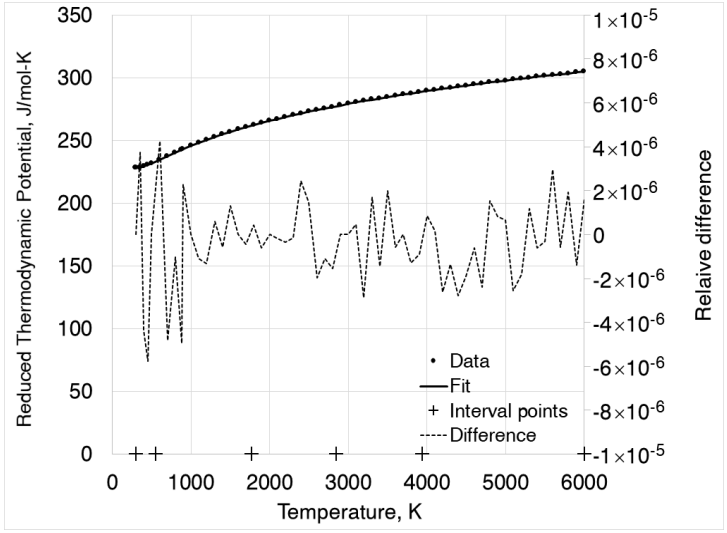
Sulfur (S_2) – S2(g)

MW = 64.12 g/mol – $\Delta_f H^o$(298.15 K) = 128.6 kJ/mol – S^o(298.15 K) = 228.165 J/mol-K – p^o = 0.1 MPa



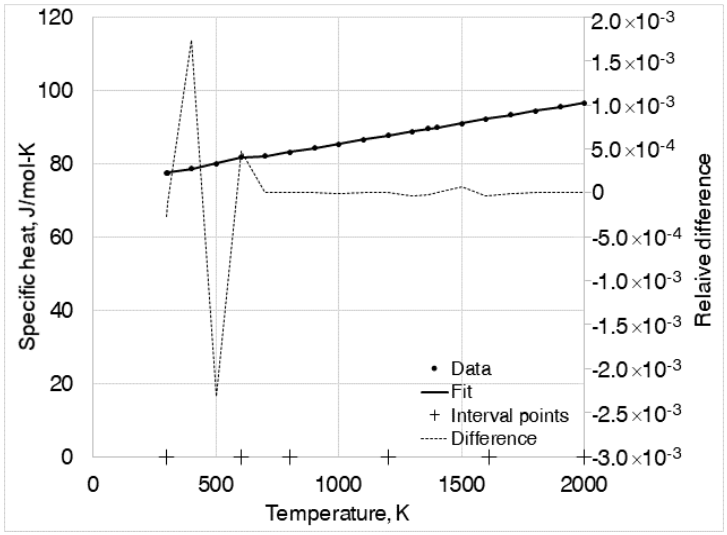


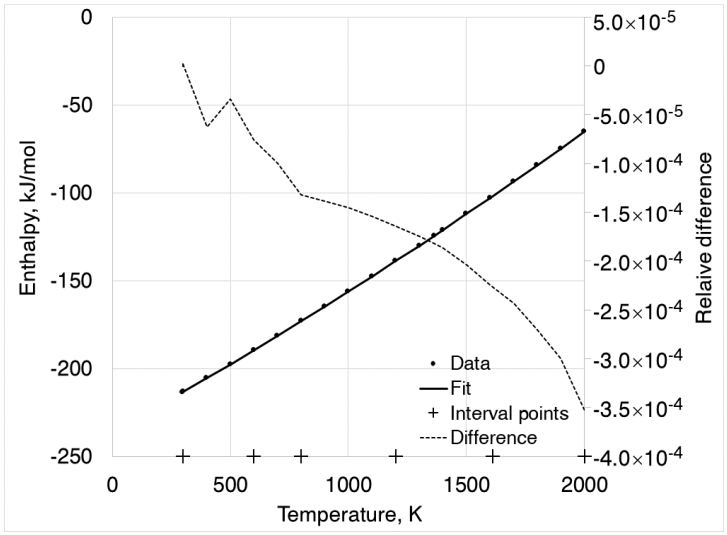


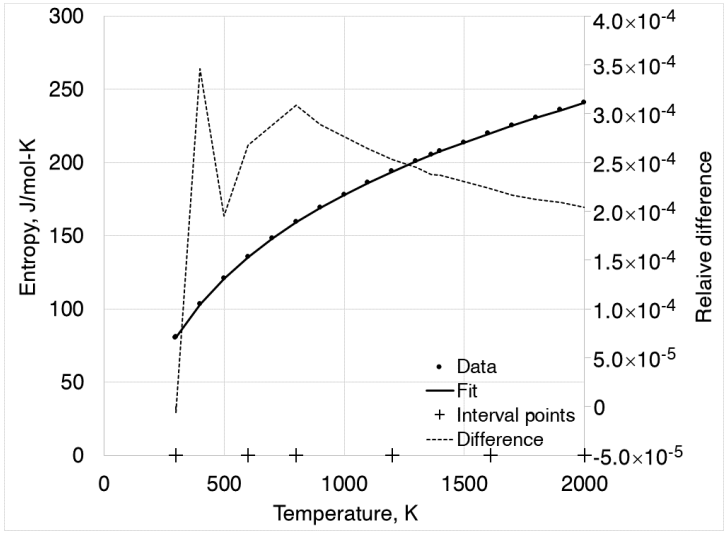


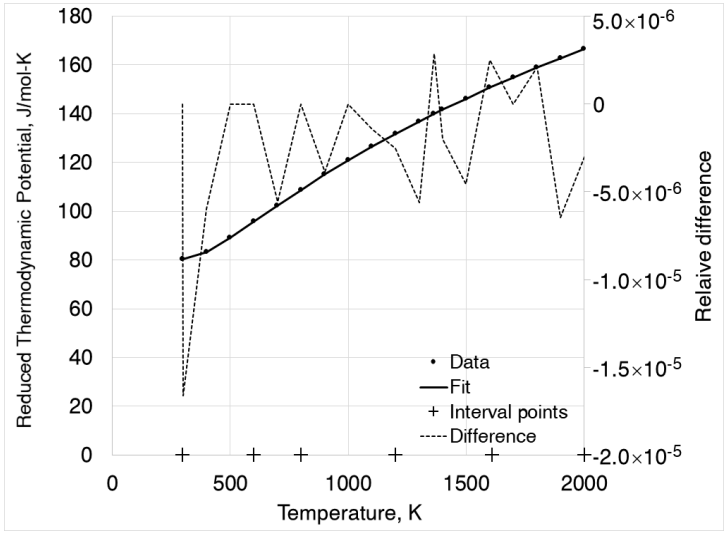
Silicon Sulfide (SiS_2) – S2Si1(cr)

MW = 92.21 g/mol – $\Delta_f H^o$(298.15 K) = -213.384 kJ/mol – S^o(298.15 K) = 80.333 J/mol-K – p^o = 0.1 MPa



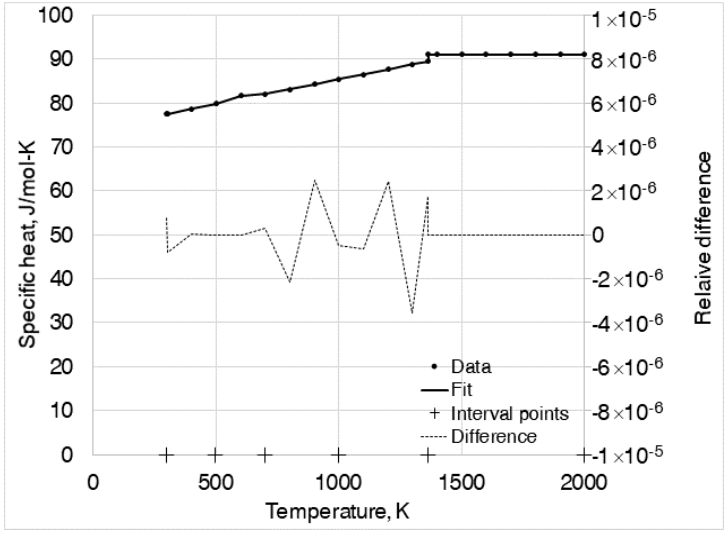


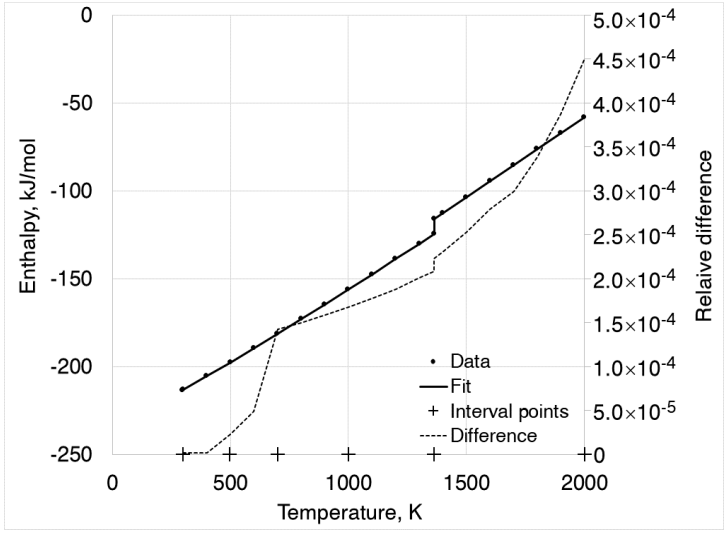


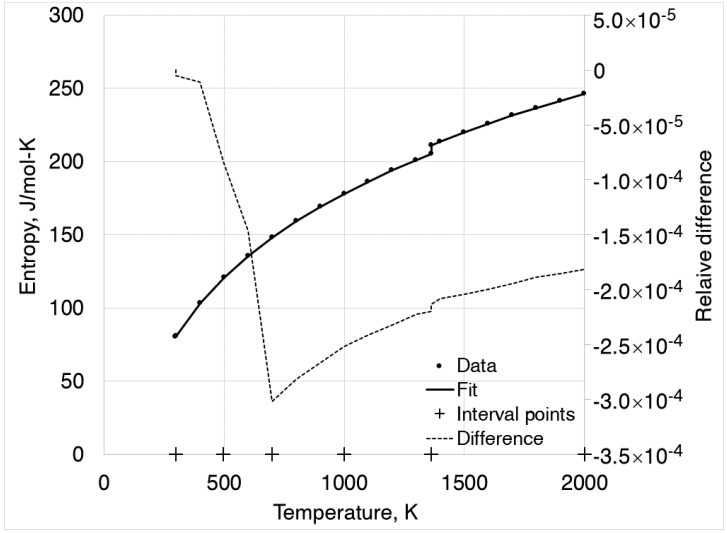


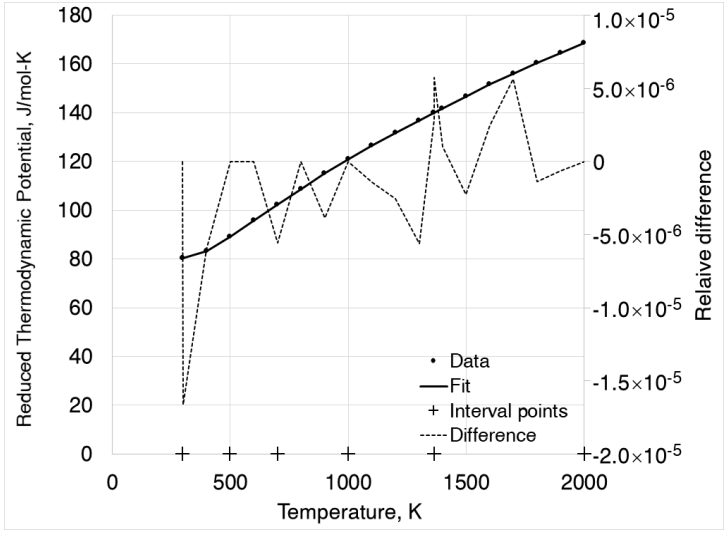
Silicon Sulfide (SiS_2) – S2Si1(cr,l)

MW = 92.21 g/mol – $\Delta_f H^o$(298.15 K) = -213.384 kJ/mol – S^o(298.15 K) = 80.333 J/mol-K – p^o = 0.1 MPa



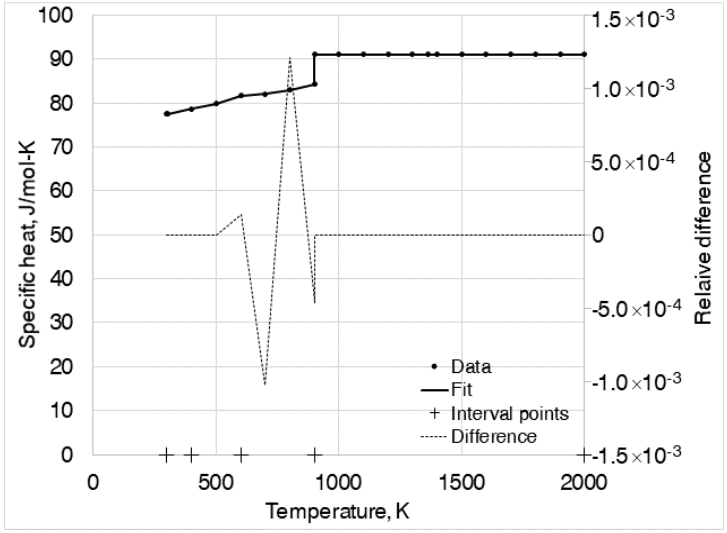


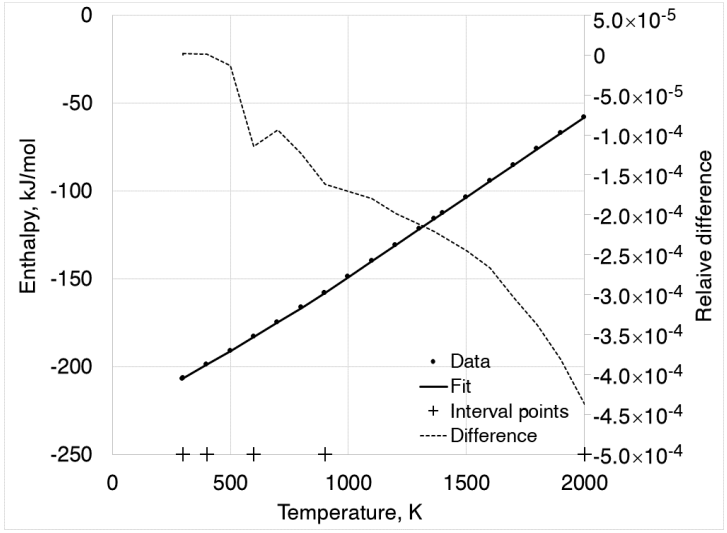


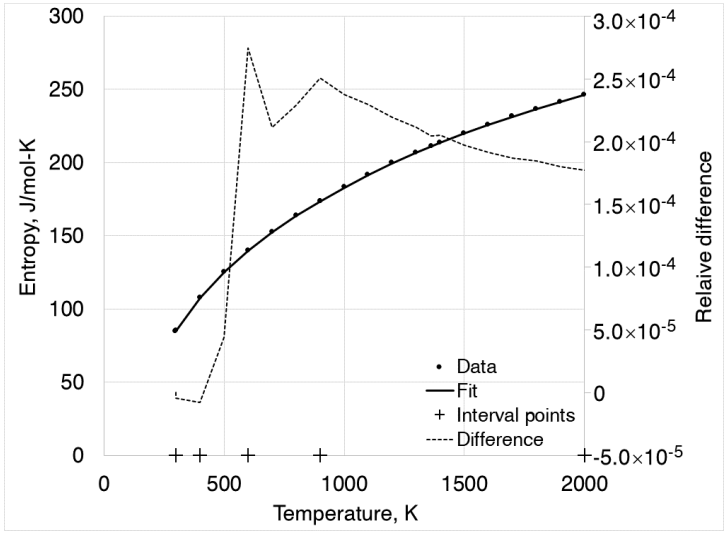


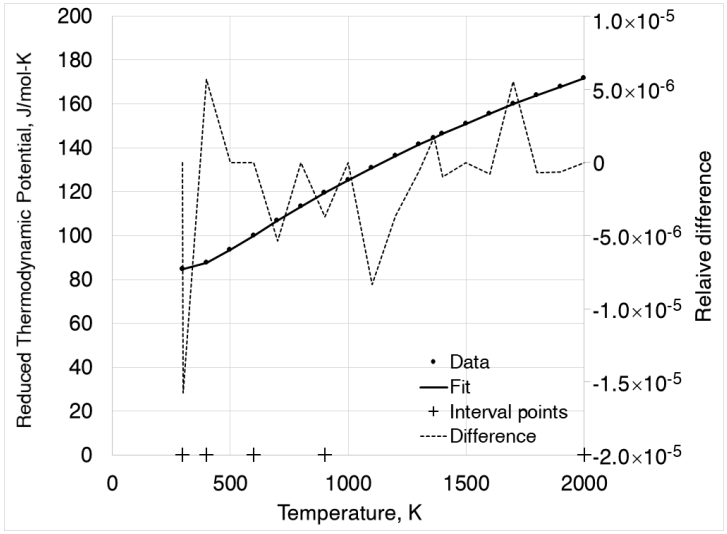
Silicon Sulfide (SiS_2) – S2Si1(l)

MW = 92.21 g/mol – $\Delta_f H^o$(298.15 K) = -206.924 kJ/mol – S^o(298.15 K) = 84.687 J/mol-K – p^o = 0.1 MPa



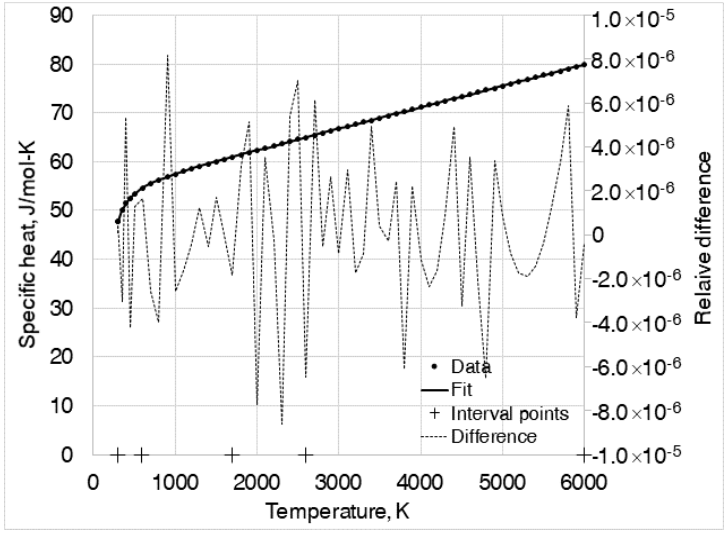


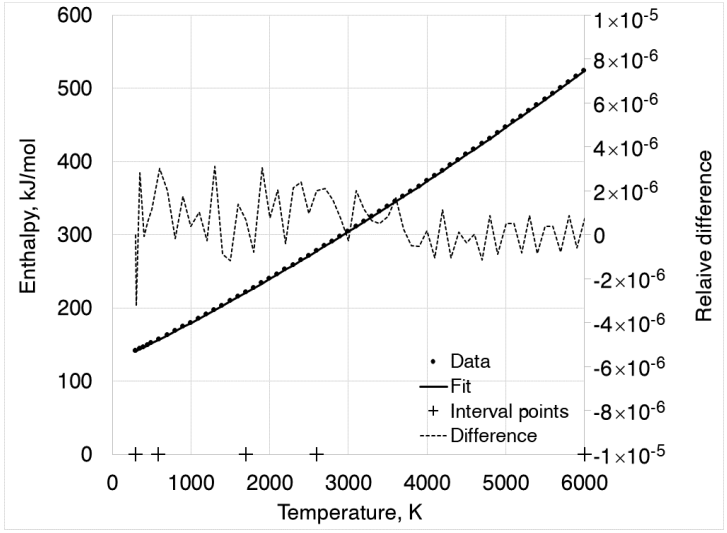


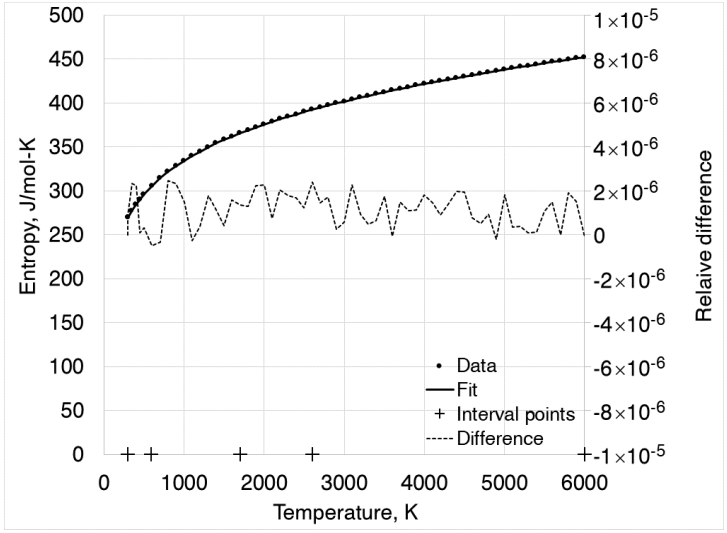


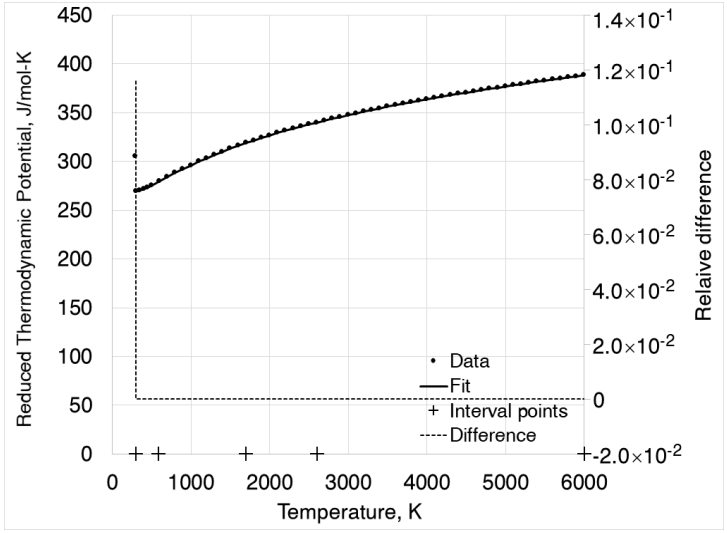
Sulfur (S_3) – S3(g)

MW = 96.18 g/mol – $\Delta_f H^o$(298.15 K) = 141.461 kJ/mol – S^o(298.15 K) = 269.517 J/mol-K – p^o = 0.1 MPa



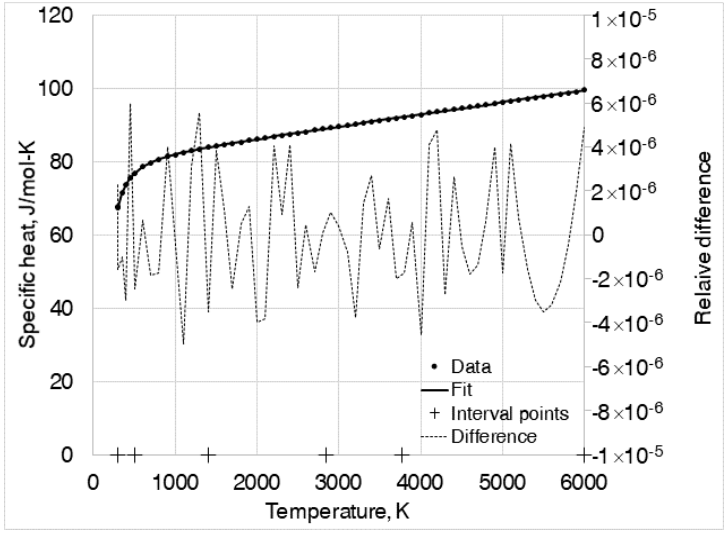


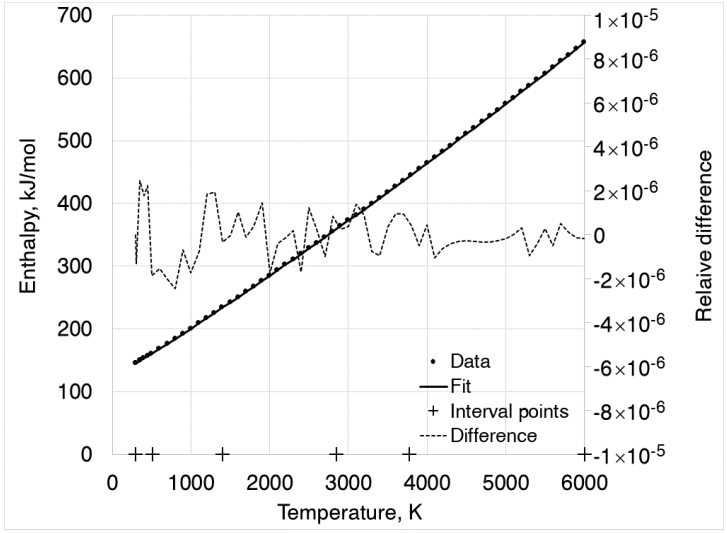


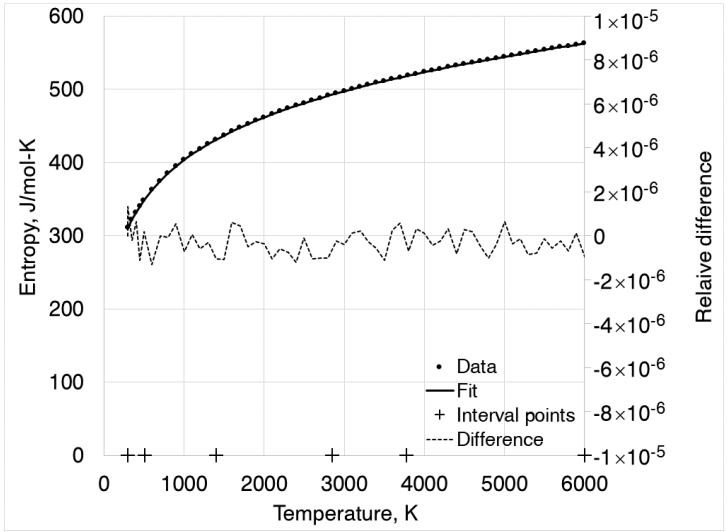


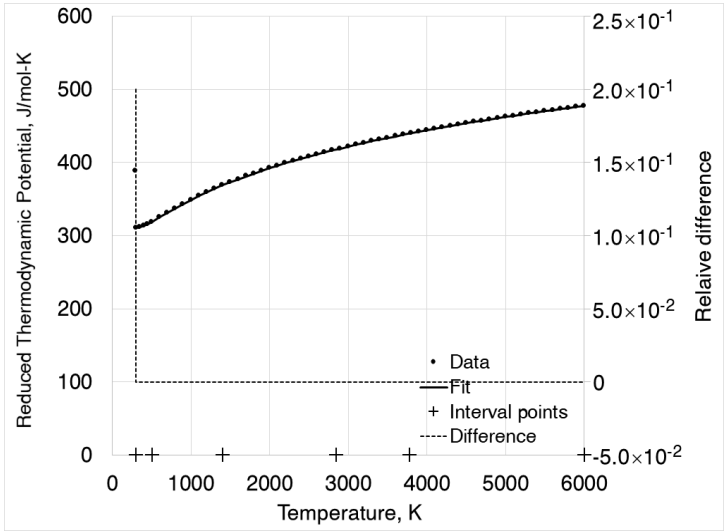
Sulfur (S_4) – S4(g)

MW = 128.24 g/mol – $\Delta_f H^o$(298.15 K) = 145.771 kJ/mol – S^o(298.15 K) = 310.646 J/mol-K – p^o = 0.1 MPa



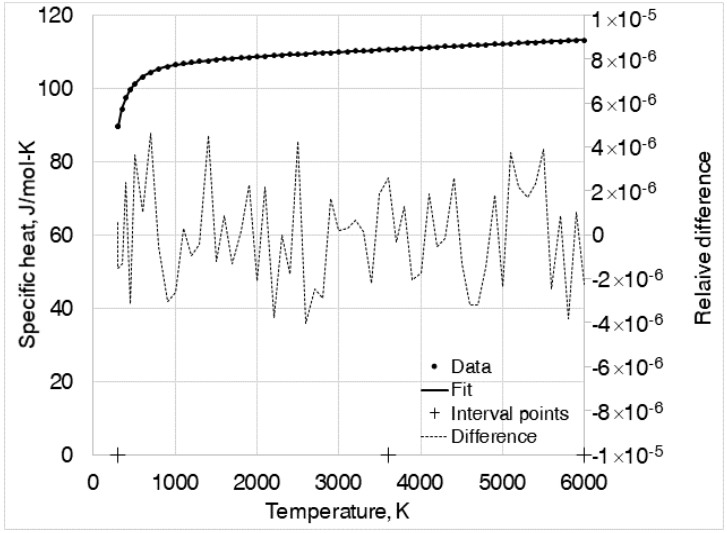


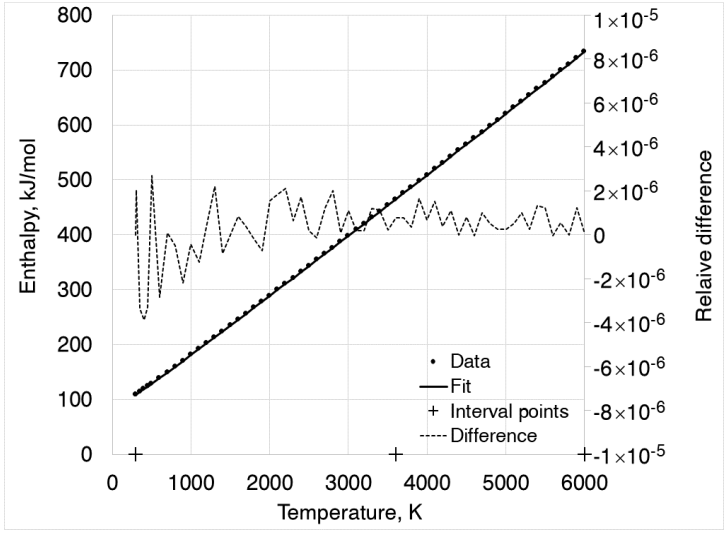


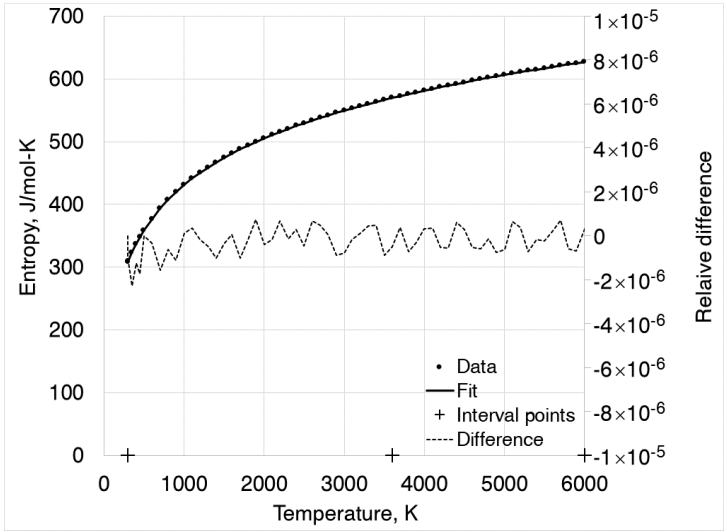


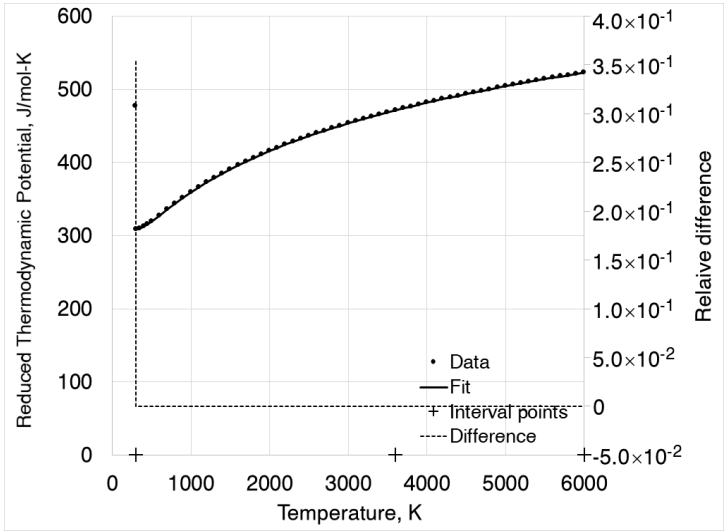
Sulfur (S_5) – S5(g)

MW = 160.3 g/mol – $\Delta_f H^o$(298.15 K) = 109.37 kJ/mol – S^o(298.15 K) = 308.638 J/mol-K – p^o = 0.1 MPa



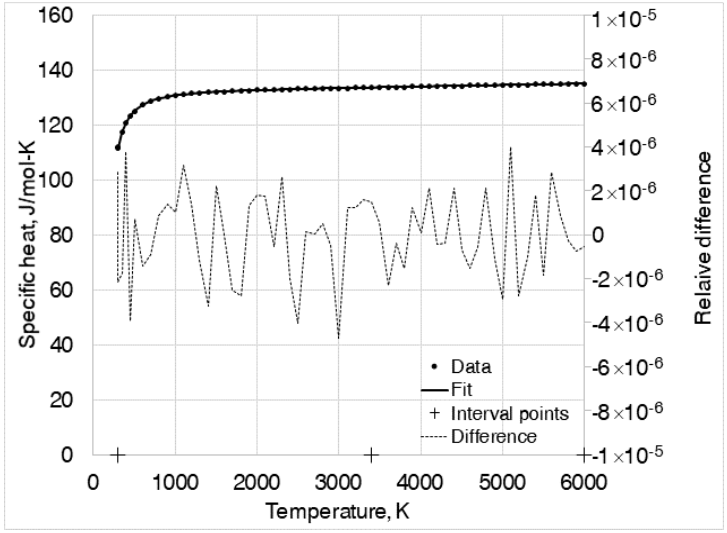


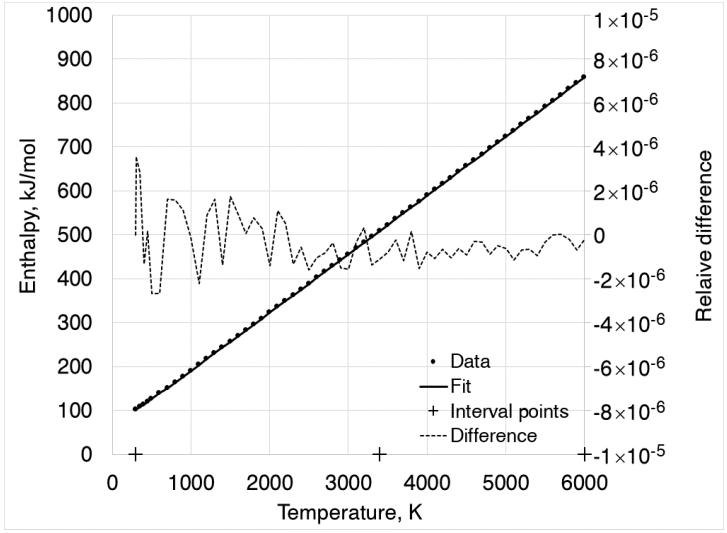


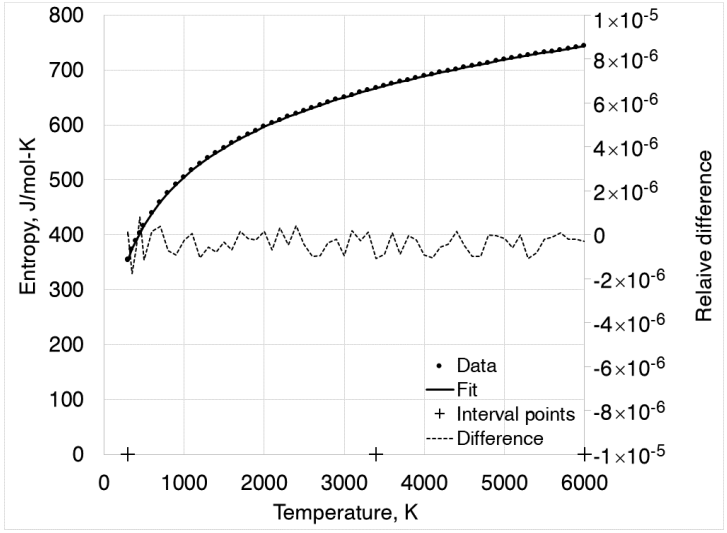


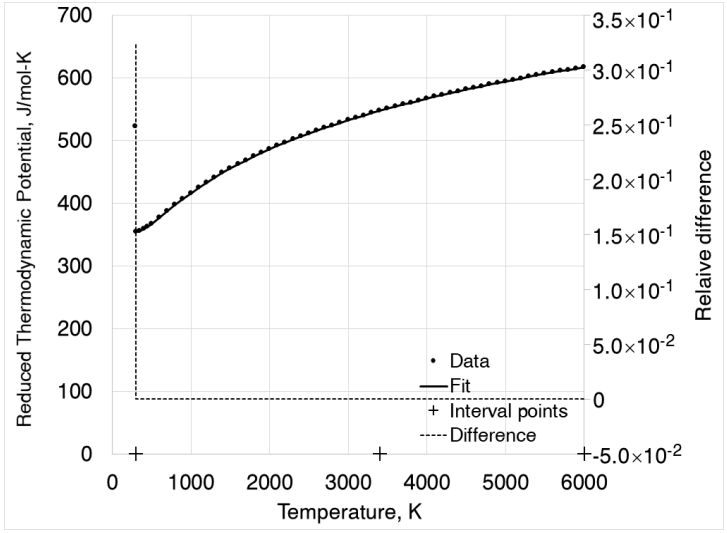
Sulfur (S_6) – S6(g)

MW = 192.36 g/mol – $\Delta_f H^o$(298.15 K) = 101.922 kJ/mol – S^o(298.15 K) = 354.076 J/mol-K – p^o = 0.1 MPa



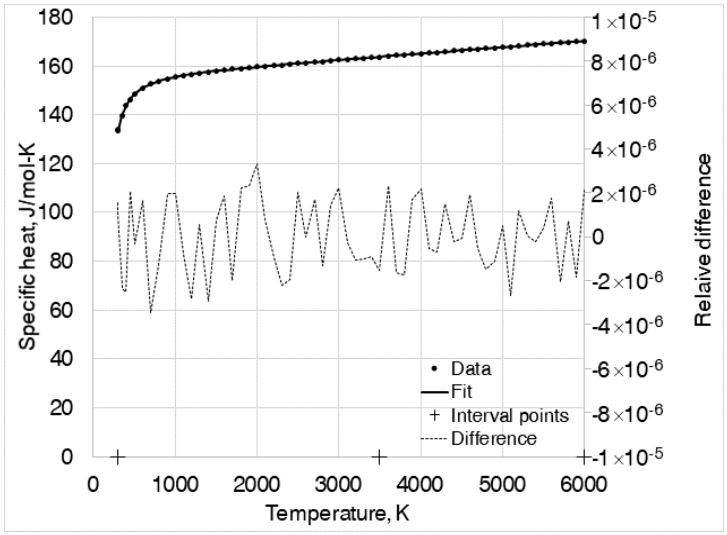


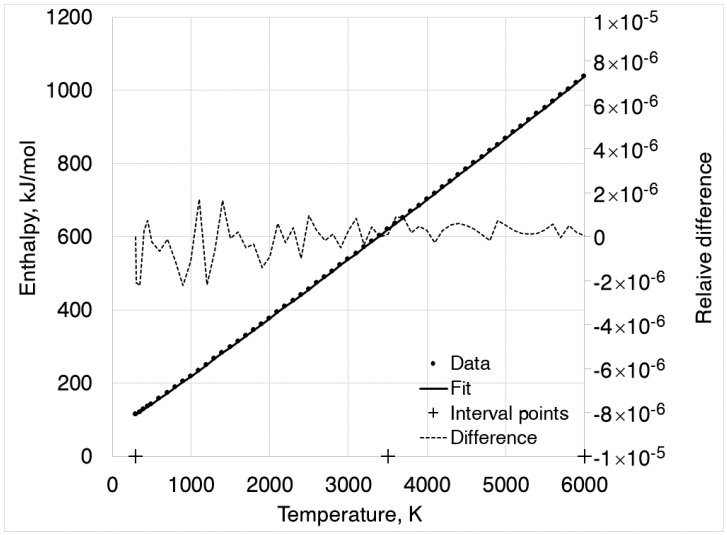


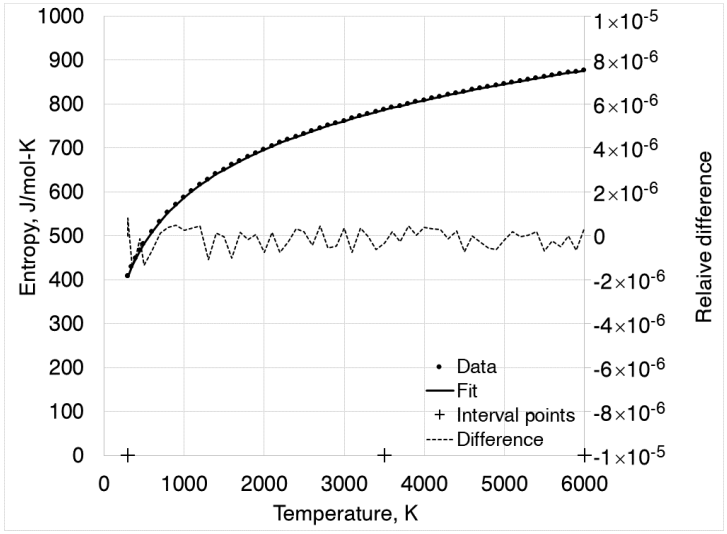


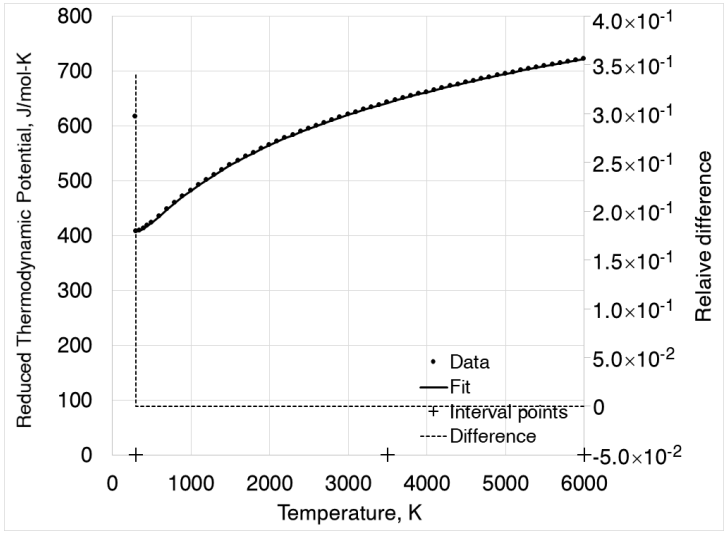
Sulfur (S_7) – S7(g)

MW = 224.42 g/mol – $\Delta_f H^o$(298.15 K) = 113.679 kJ/mol – S^o(298.15 K) = 407.673 J/mol-K – p^o = 0.1 MPa



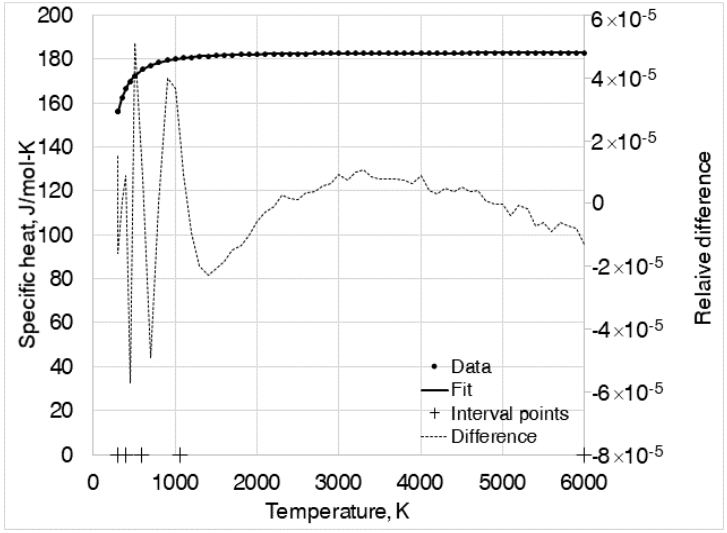


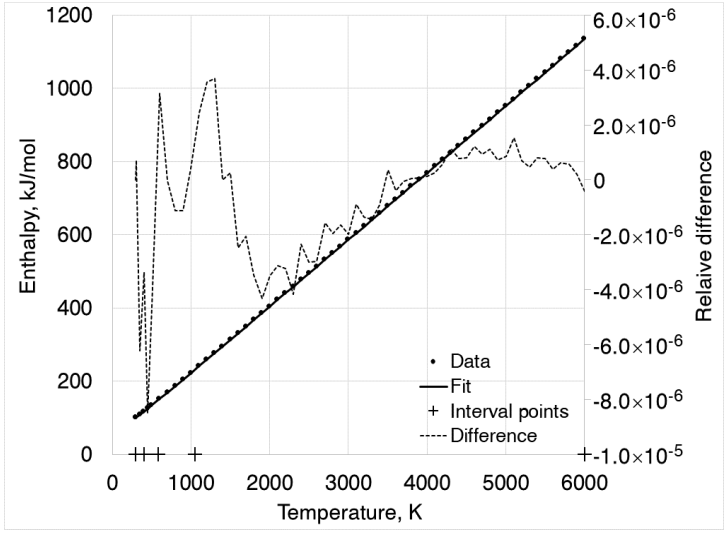


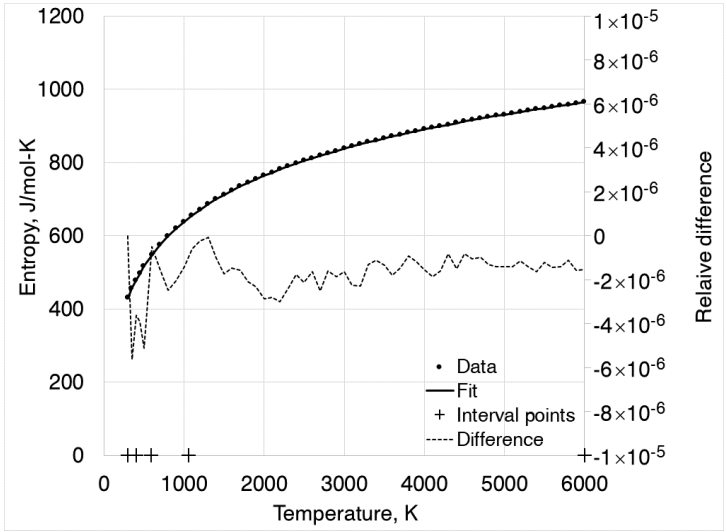


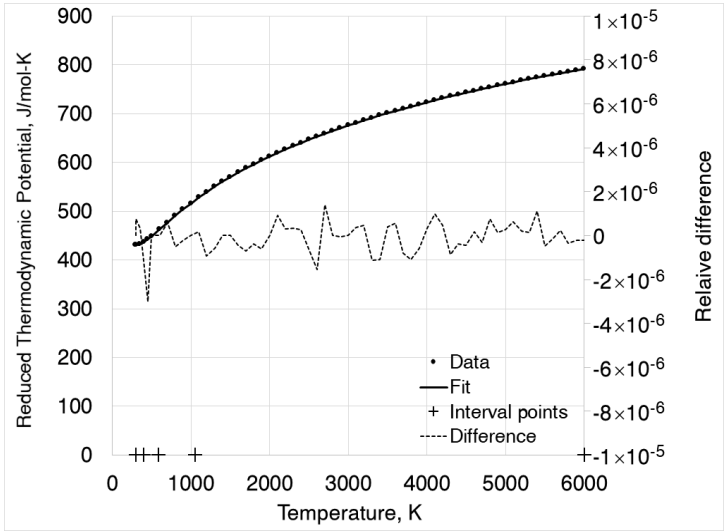
Sulfur (S_8) – S8(g)

MW = 256.48 g/mol – $\Delta_f H^o$(298.15 K) = 100.416 kJ/mol – S^o(298.15 K) = 430.311 J/mol-K – p^o = 0.1 MPa



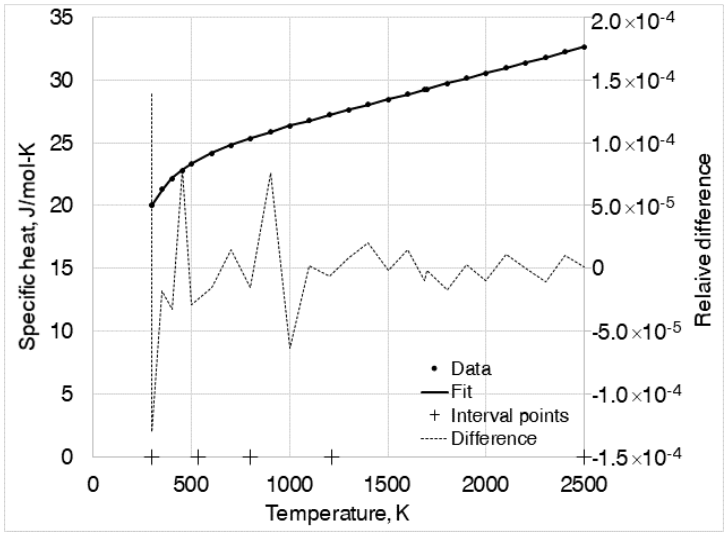


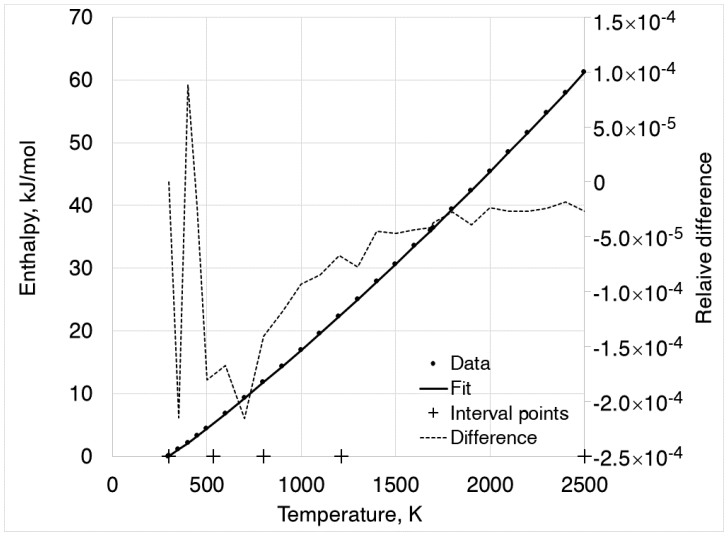


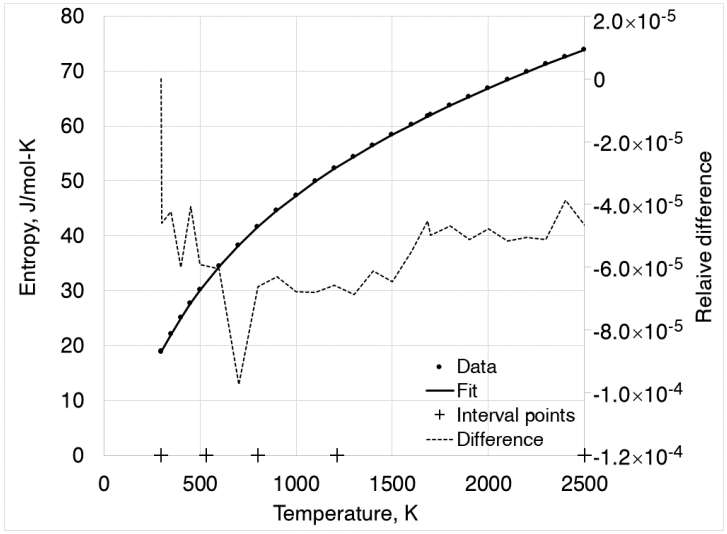


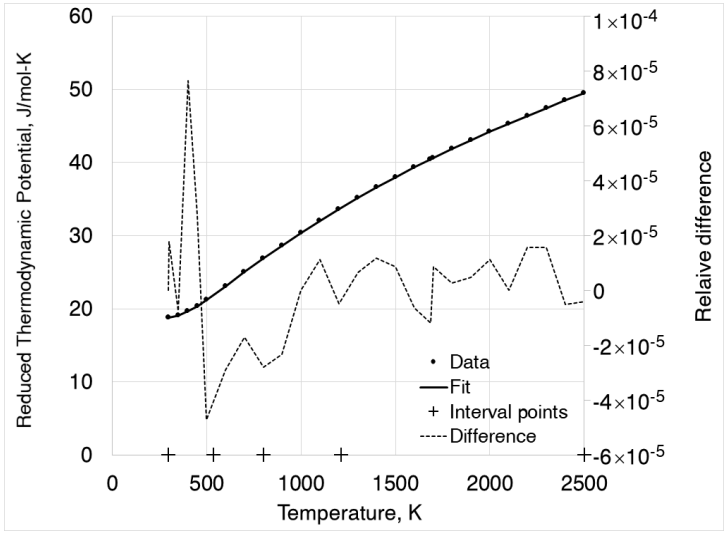
Silicon (Si) – Si1(cr)

MW = 28.0855 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 18.82 J/mol-K – p^o = 0.1 MPa



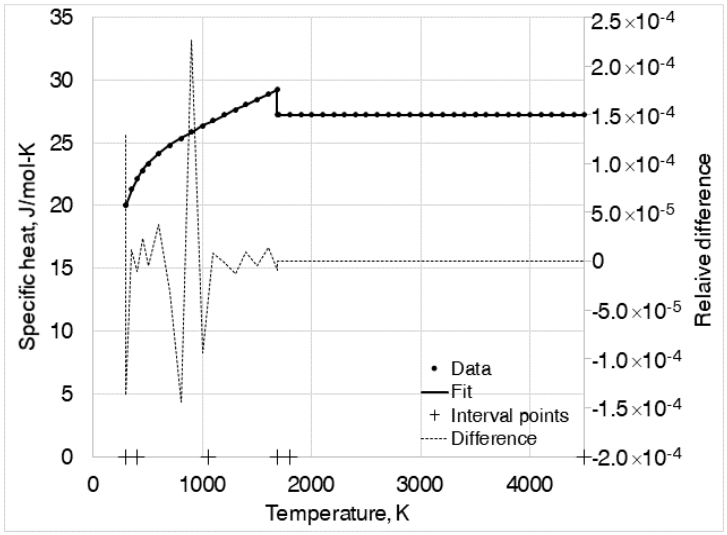


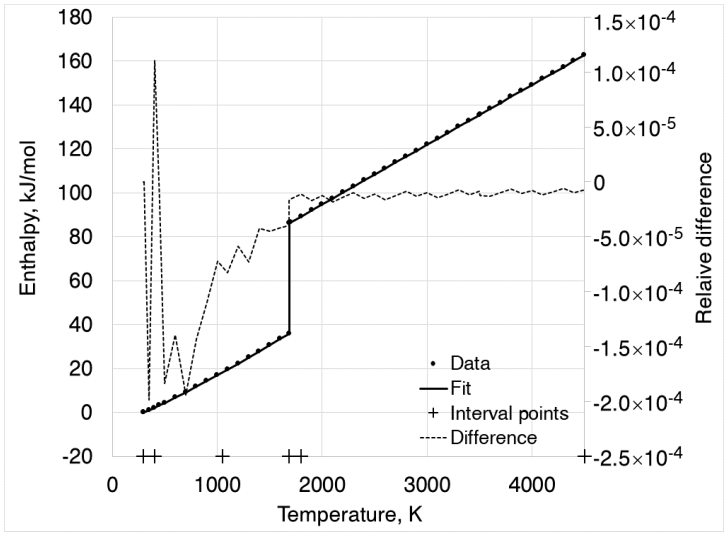


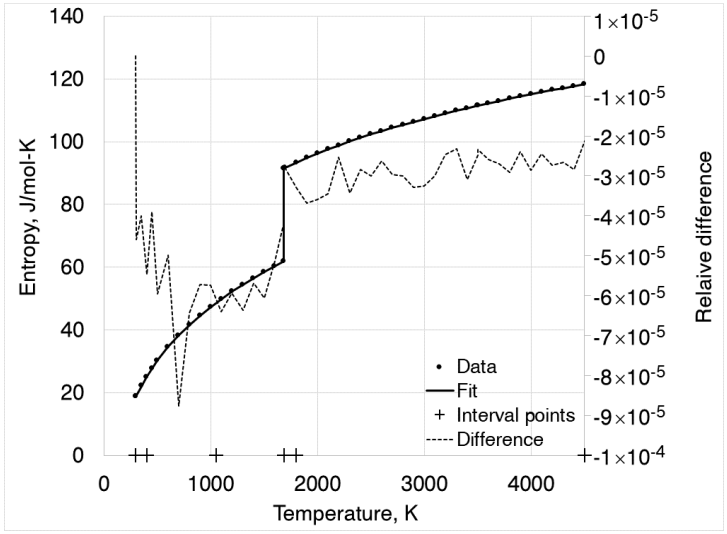


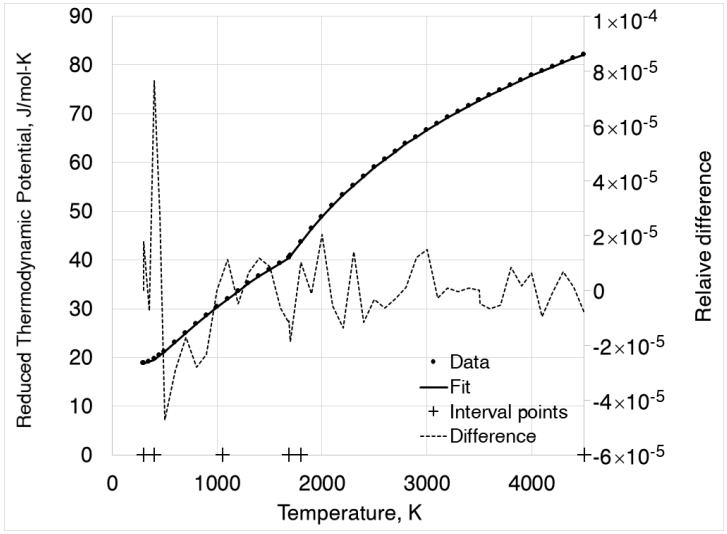
Silicon (Si) – Si1(cr,l)

MW = 28.0855 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 18.82 J/mol-K – p^o = 0.1 MPa



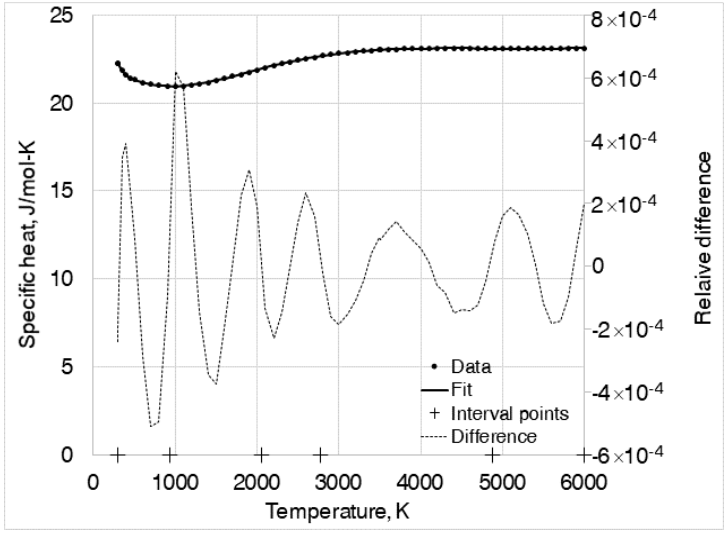


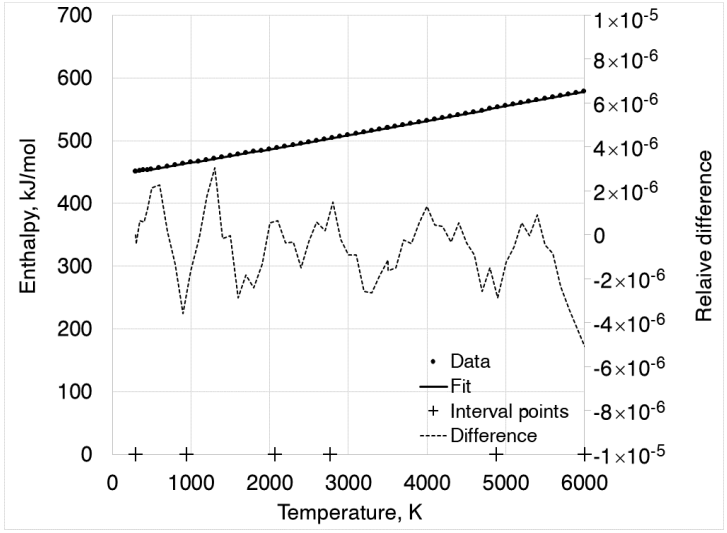


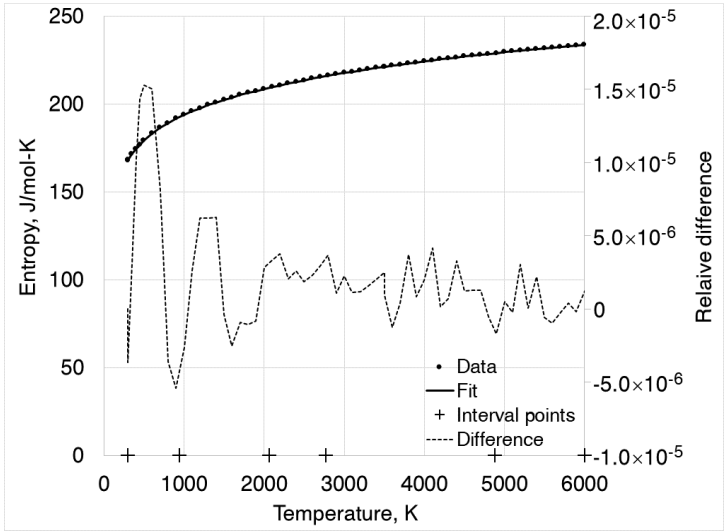


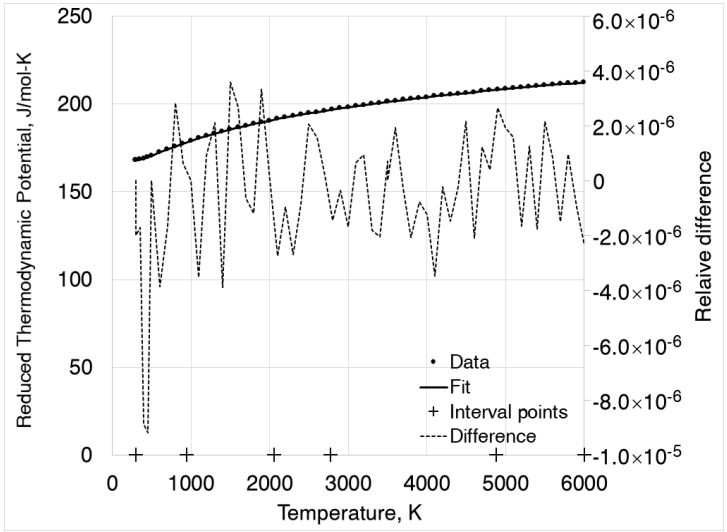
Silicon (Si) – Si1(g)

MW = 28.0855 g/mol – $\Delta_f H^o$(298.15 K) = 450 kJ/mol – S^o(298.15 K) = 167.98 J/mol-K – p^o = 0.1 MPa



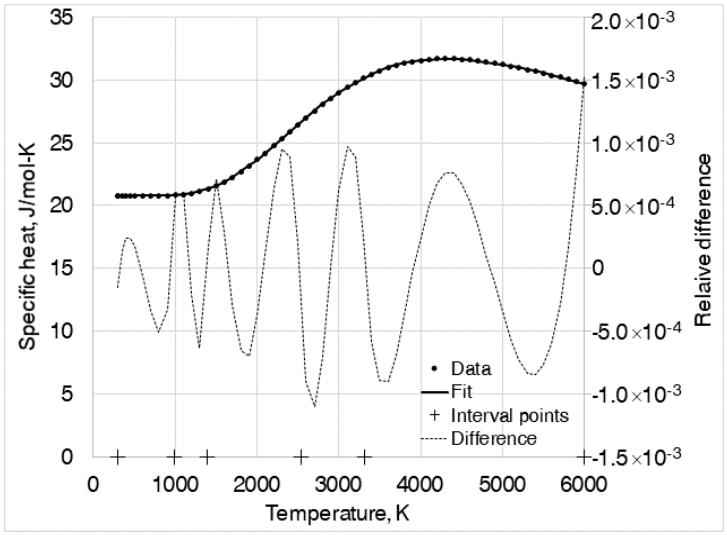


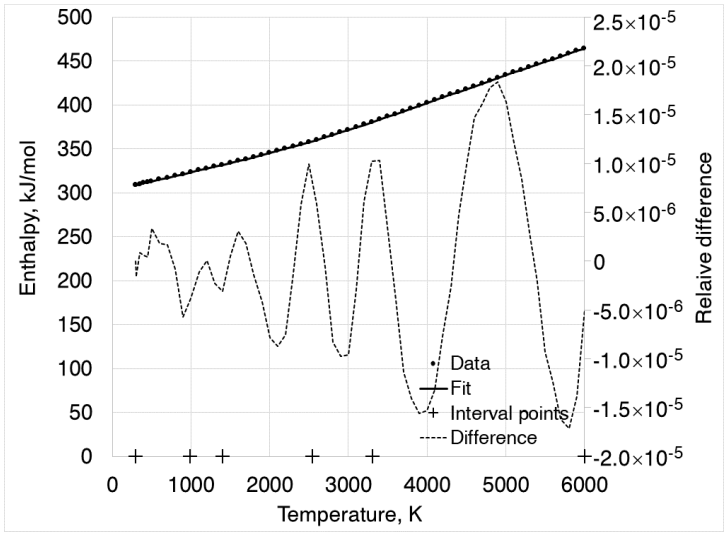


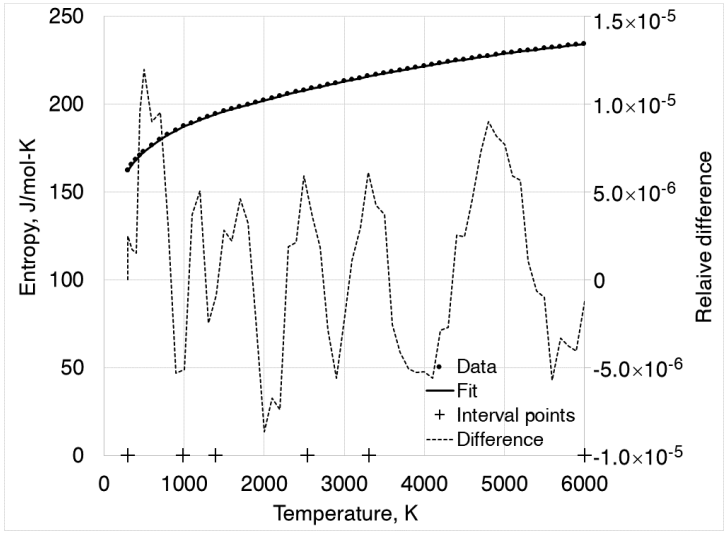


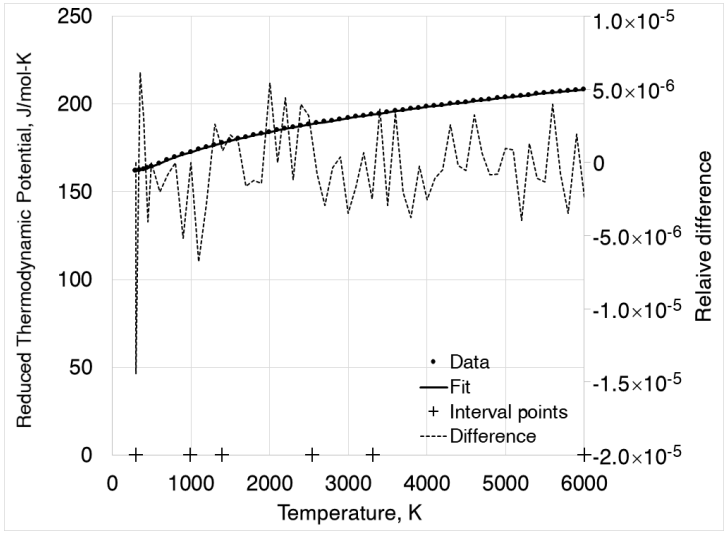
Silicon, Ion (Si^-) – Si1-(g)

MW = 28.086 g/mol – $\Delta_f H^o$(298.15 K) = 308.819 kJ/mol – S^o(298.15 K) = 161.977 J/mol-K – p^o = 0.1 MPa



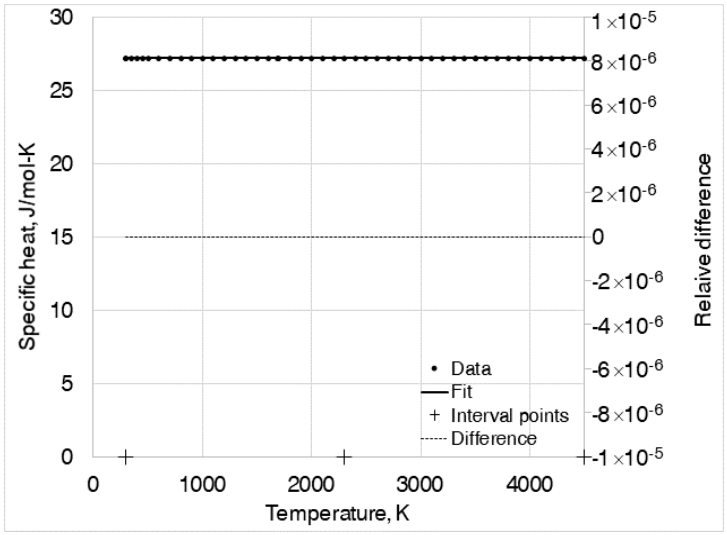


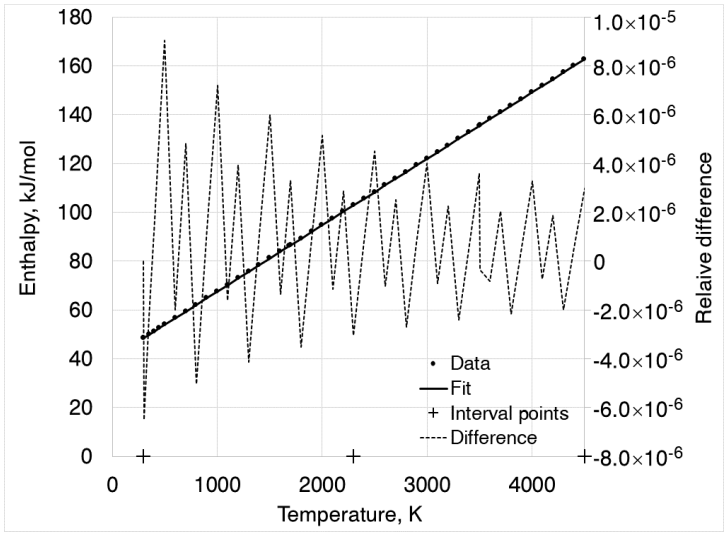


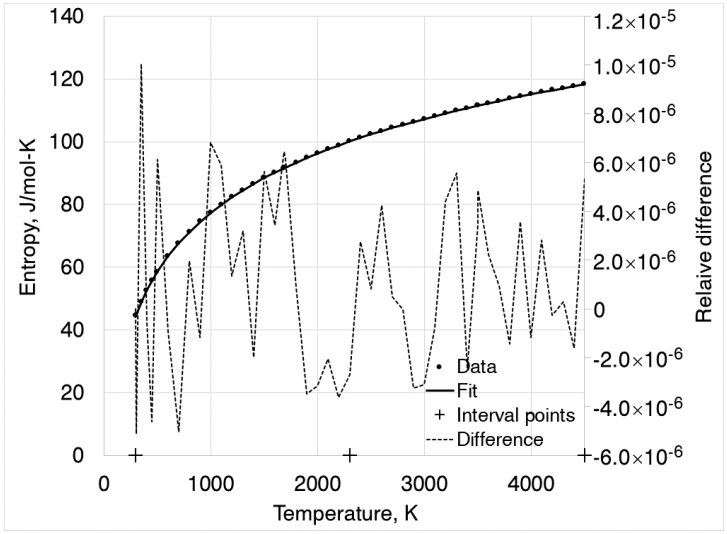


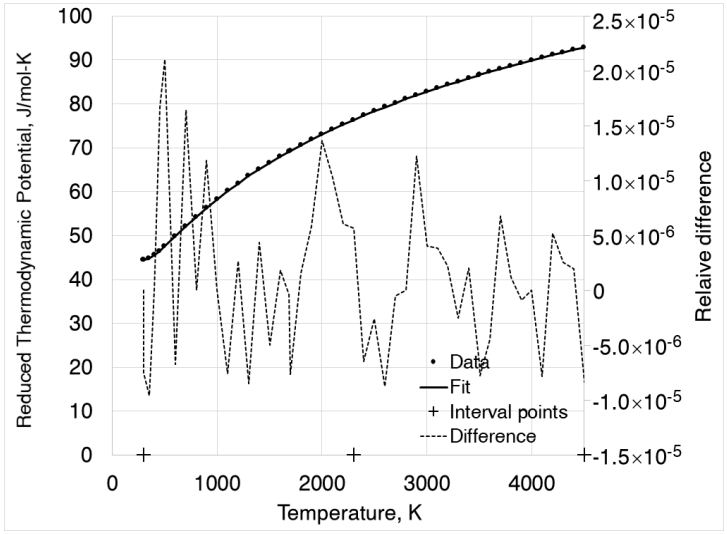
Silicon (Si) – Si1(l)

MW = 28.0855 g/mol – $\Delta_f H^o$(298.15 K) = 48.47 kJ/mol – S^o(298.15 K) = 44.46 J/mol-K – p^o = 0.1 MPa



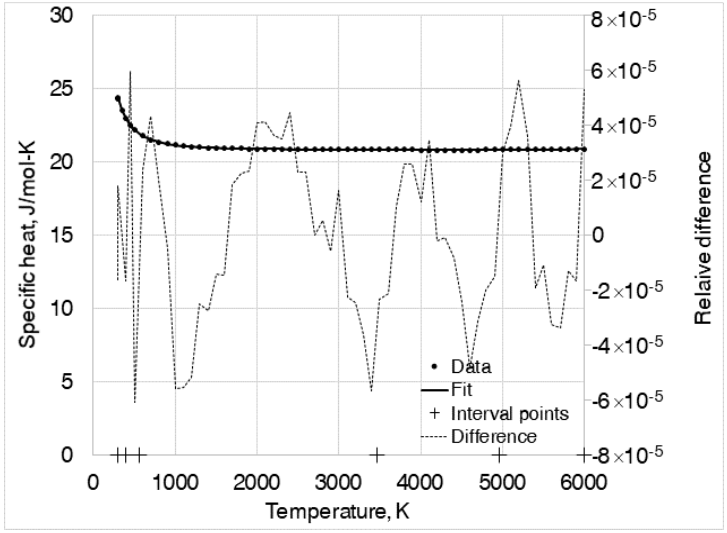


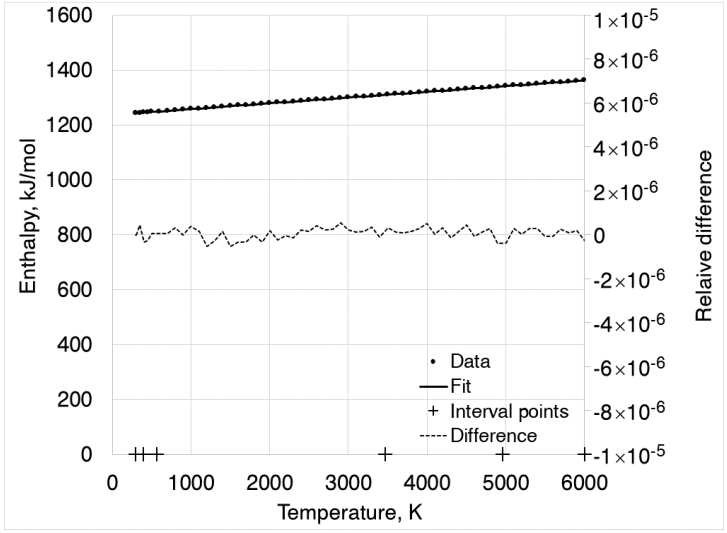


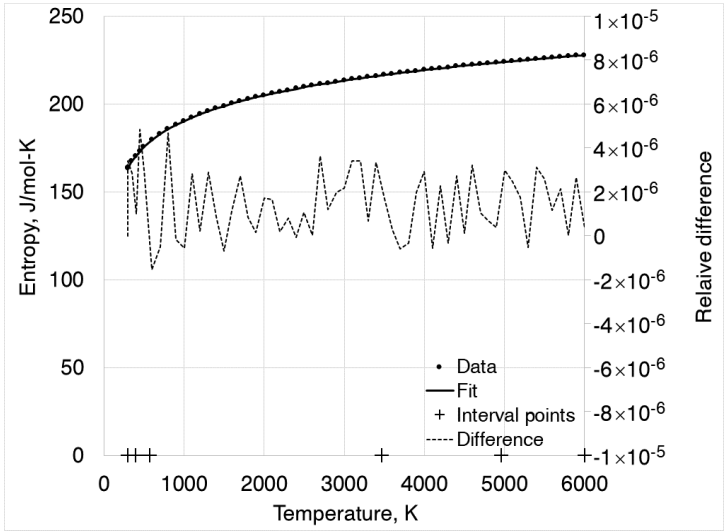


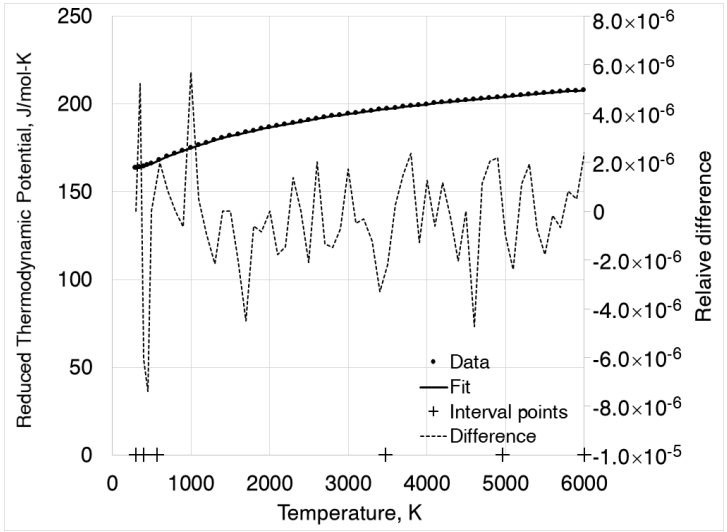
Silicon, Ion (Si^+) – Si1+(g)

MW = 28.085 g/mol – $\Delta_f H^o$(298.15 K) = 1242.505 kJ/mol – S^o(298.15 K) = 163.426 J/mol-K – p^o = 0.1 MPa



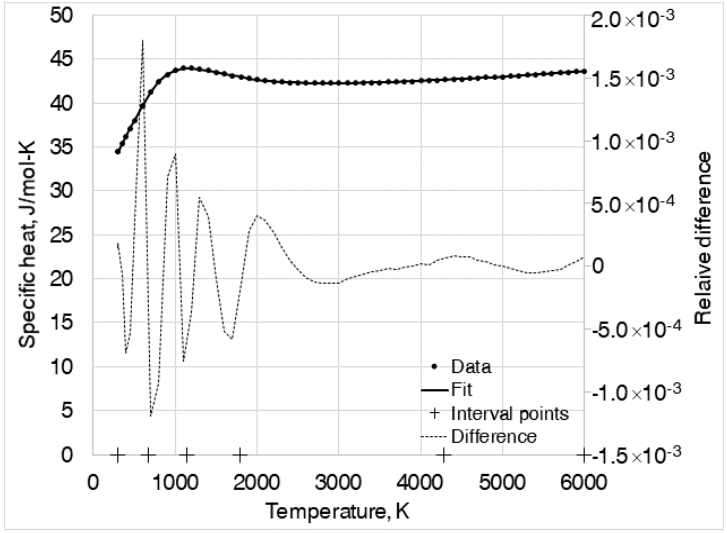


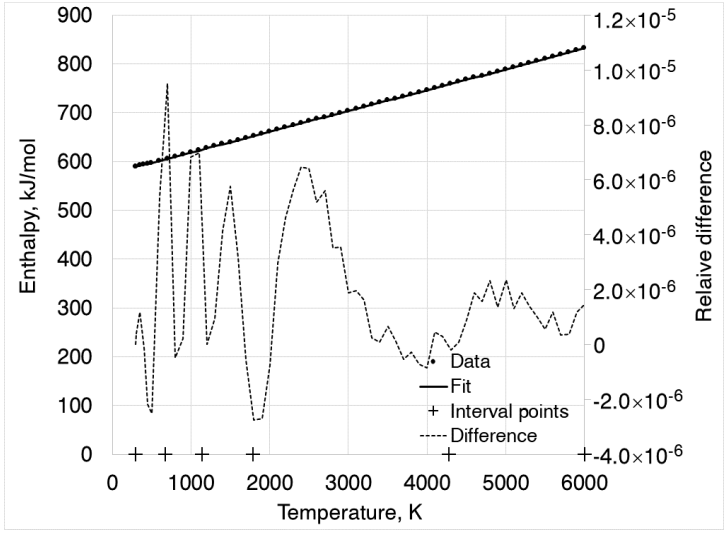


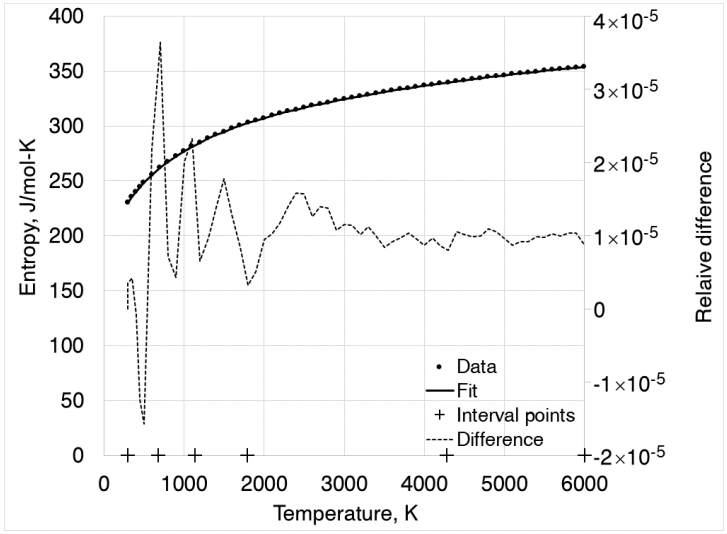


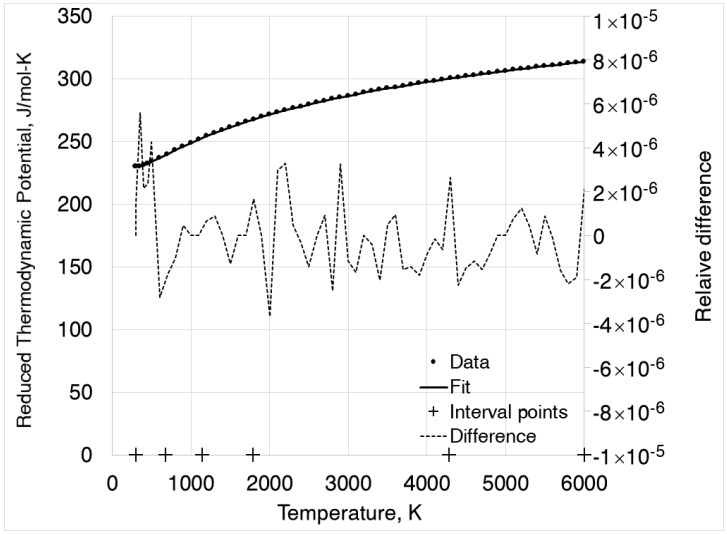
Silicon (Si_2) – Si2(g)

MW = 56.171 g/mol – $\Delta_f H^o$(298.15 K) = 589.944 kJ/mol – S^o(298.15 K) = 229.793 J/mol-K – p^o = 0.1 MPa



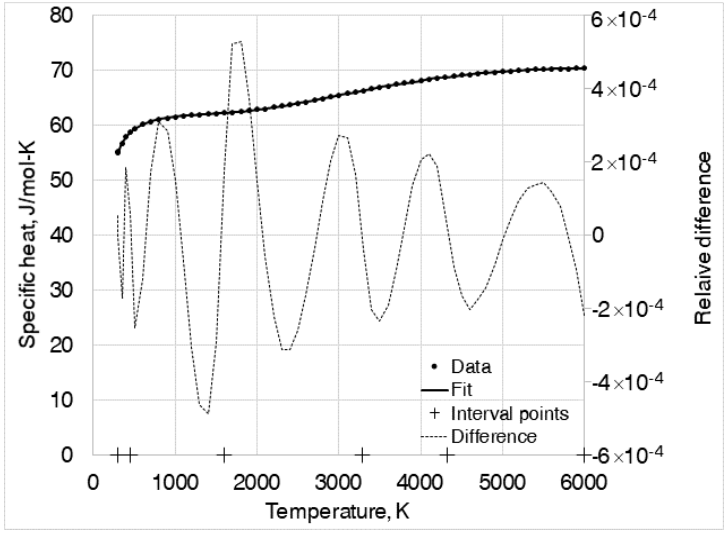


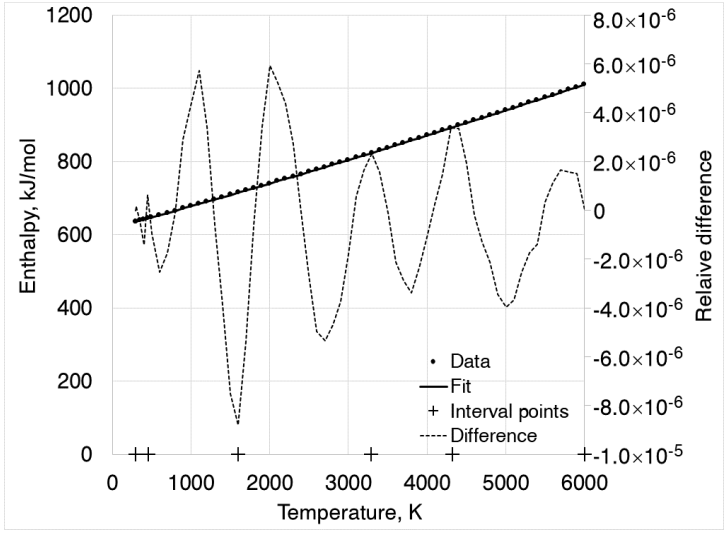


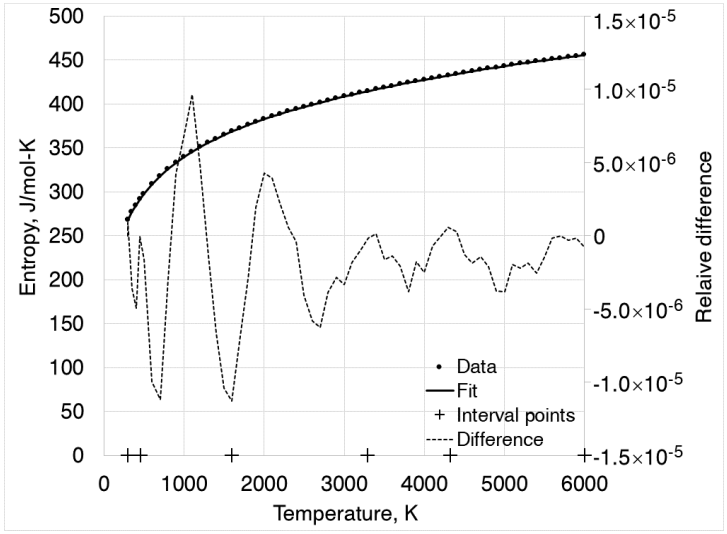


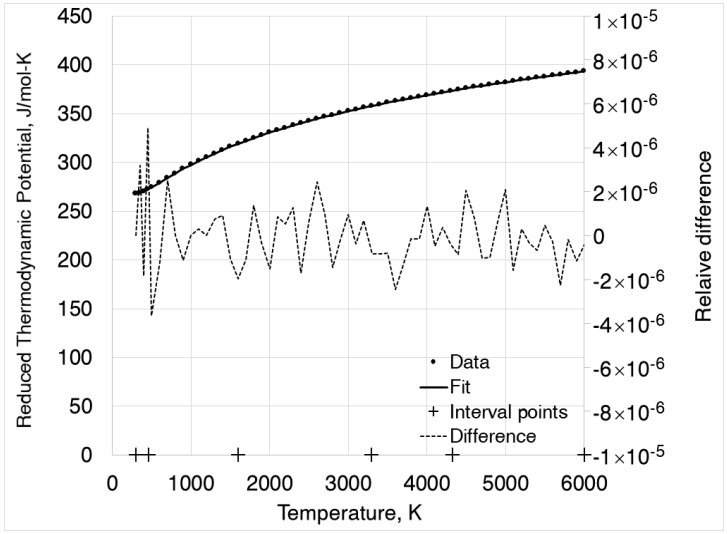
Silicon (Si_3) – Si3(g)

MW = 84.2565 g/mol – $\Delta_f H^o$(298.15 K) = 635.968 kJ/mol – S^o(298.15 K) = 267.898 J/mol-K – p^o = 0.1 MPa



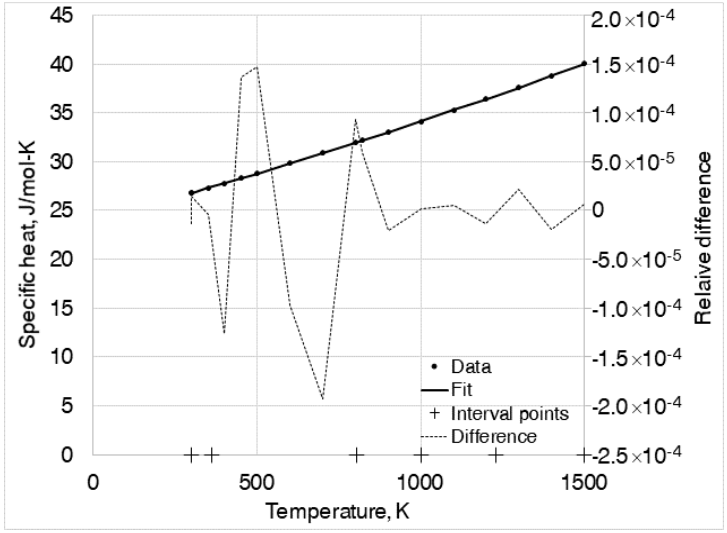


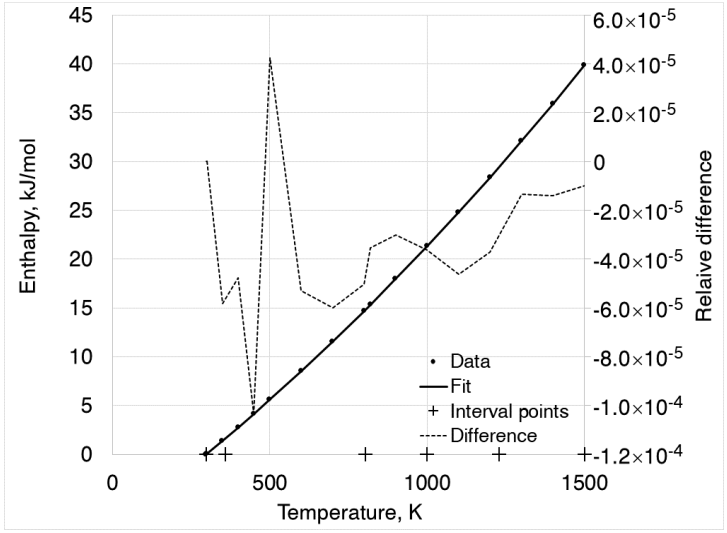


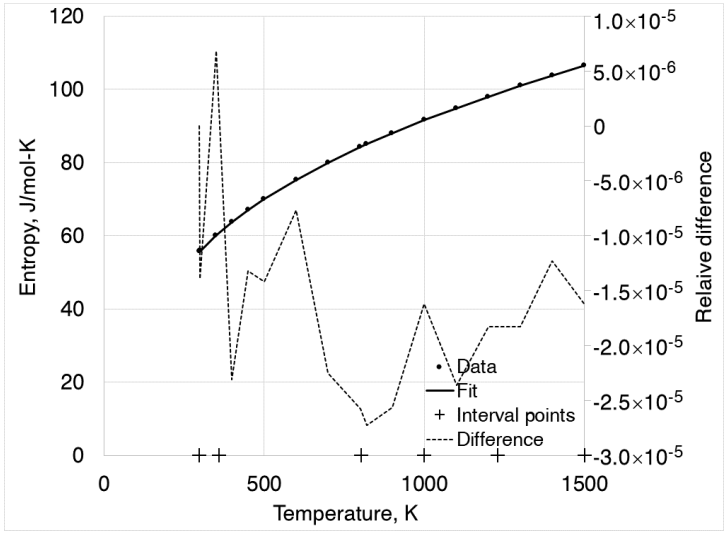


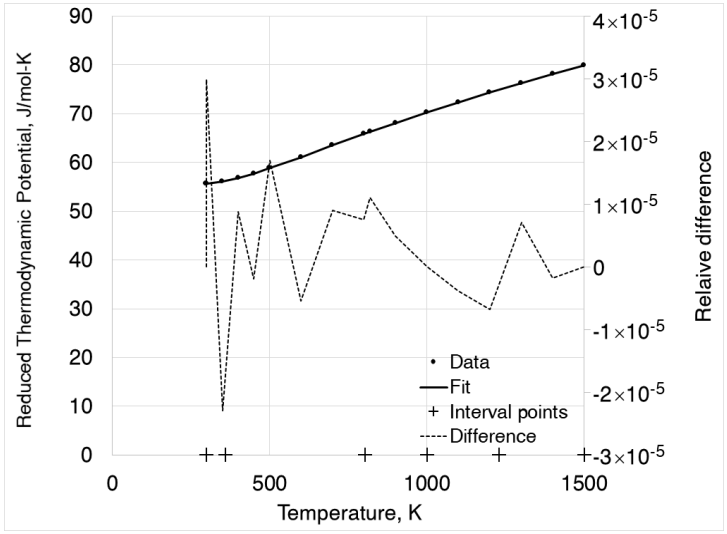
Strontium, Alpha (Sr) – Sr1(cr)_01

MW = 87.62 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 55.694 J/mol-K – p^o = 0.1 MPa



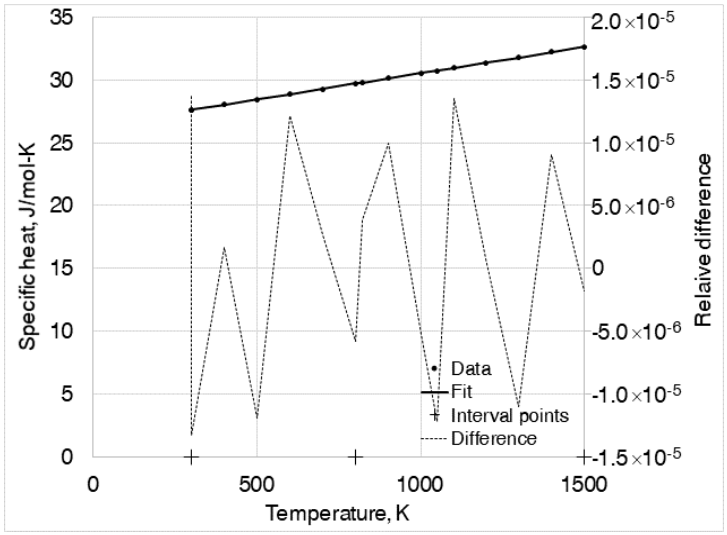


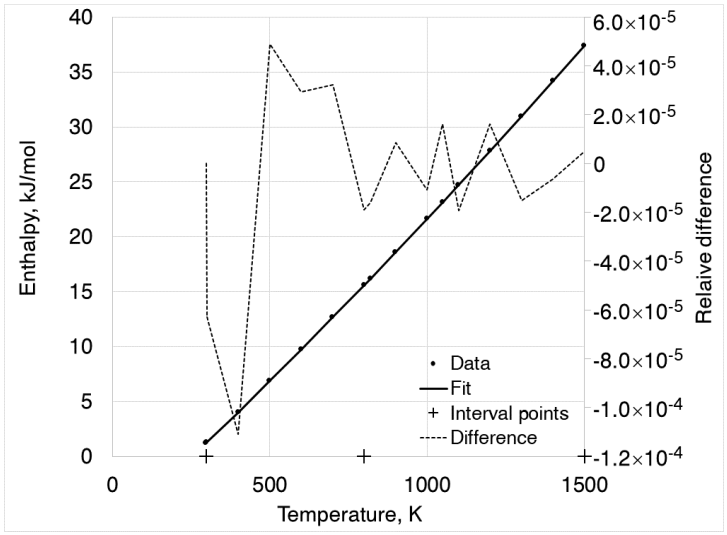


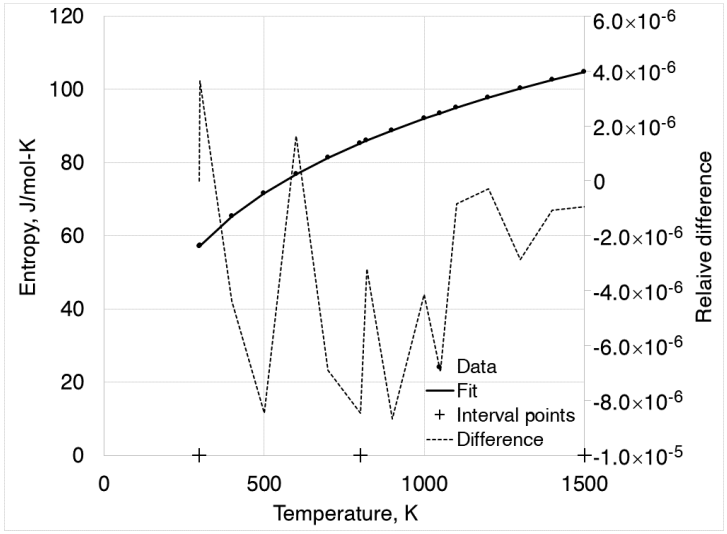


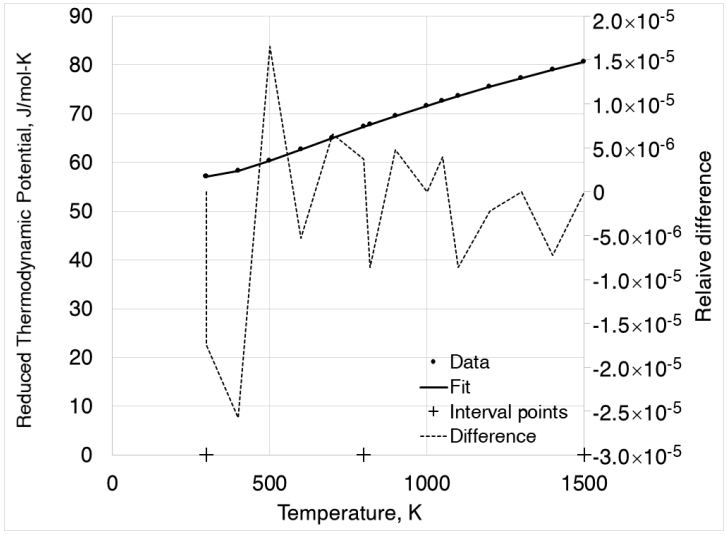
Strontium, Beta (Sr) – Sr1(cr)_02

MW = 87.62 g/mol – $\Delta_f H^o$(298.15 K) = 1.211 kJ/mol – S^o(298.15 K) = 57.175 J/mol-K – p^o = 0.1 MPa



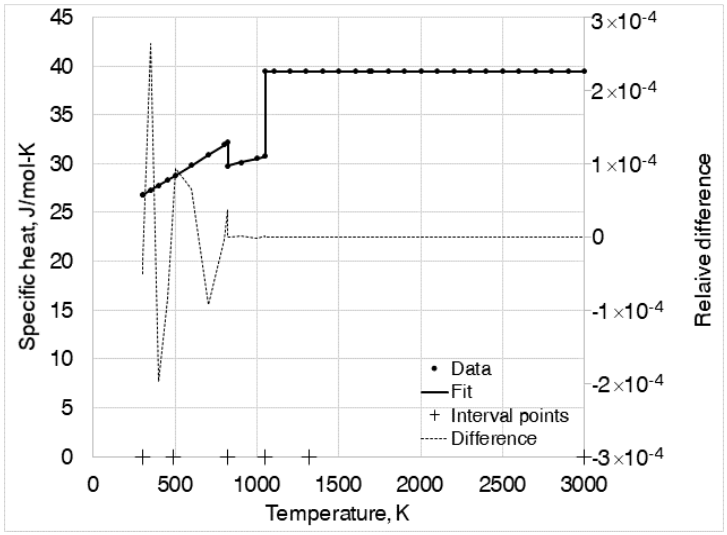


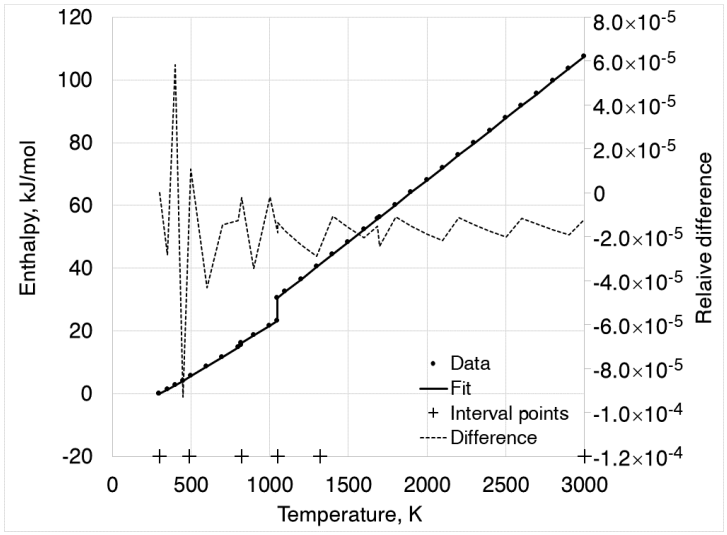


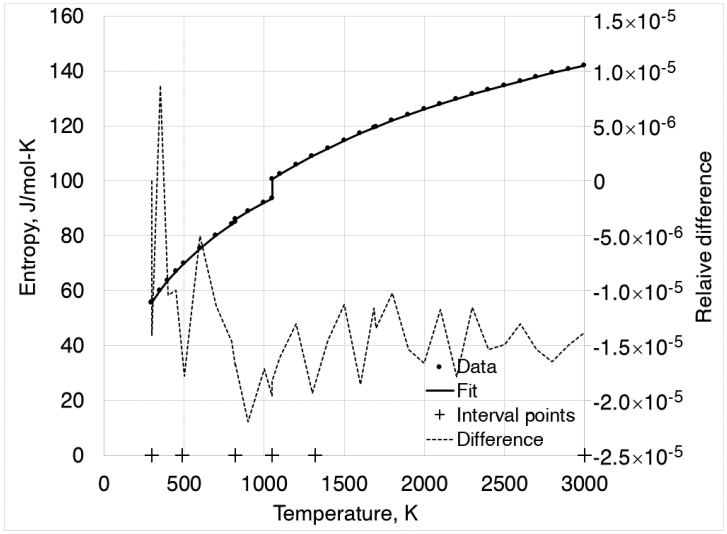


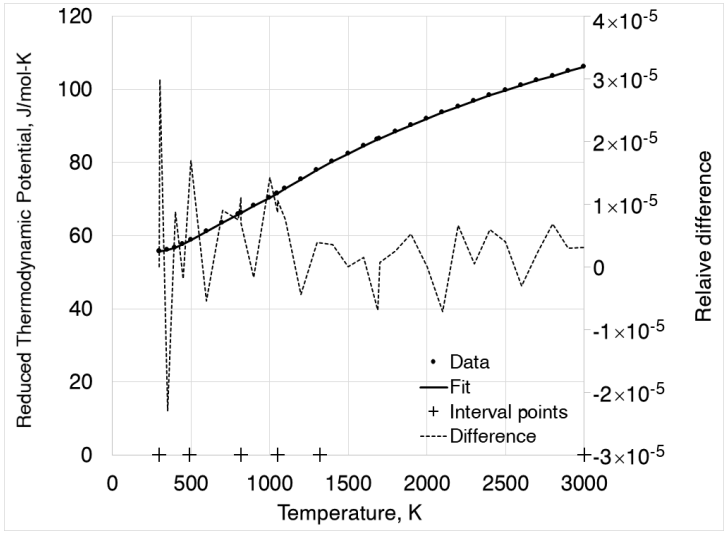
Strontium (Sr) – Sr1(cr,l)

MW = 87.62 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 55.694 J/mol-K – p^o = 0.1 MPa



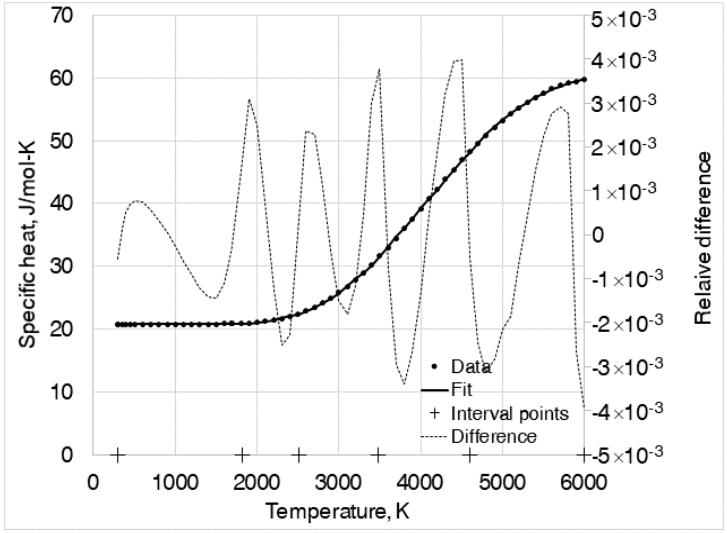


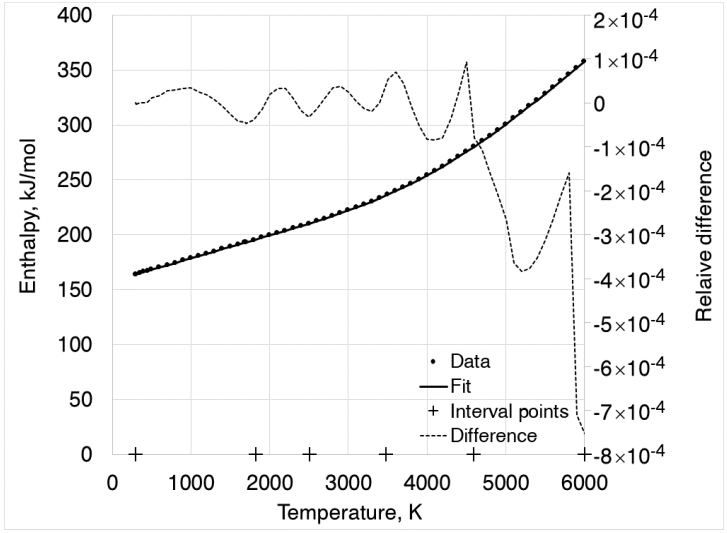


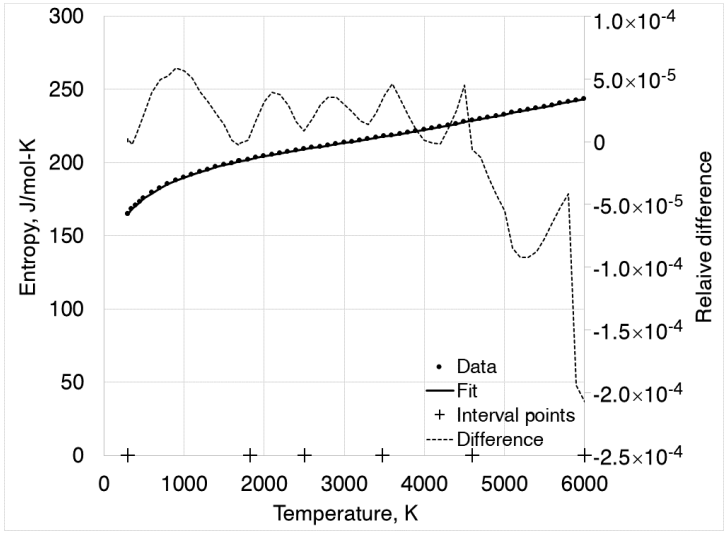


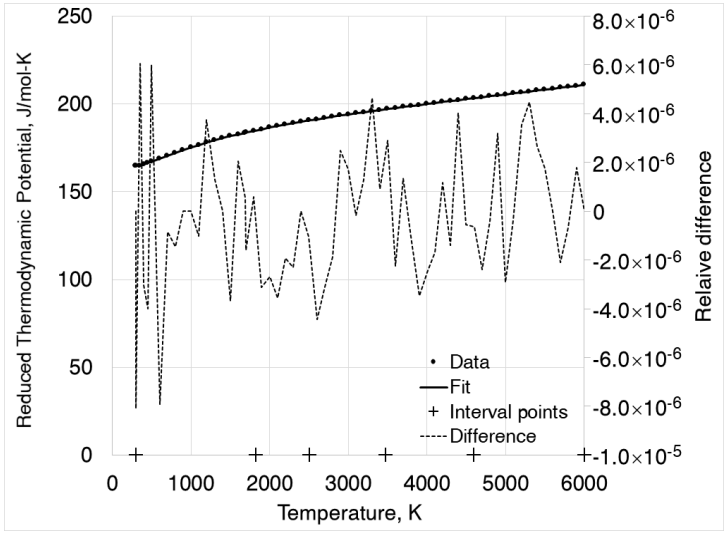
Strontium (Sr) – Sr1(g)

MW = 87.62 g/mol – $\Delta_f H^o$(298.15 K) = 164 kJ/mol – S^o(298.15 K) = 164.64 J/mol-K – p^o = 0.1 MPa



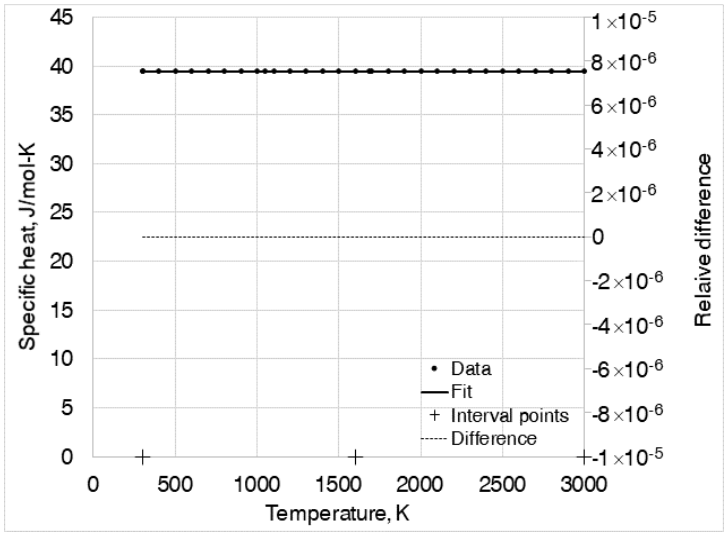


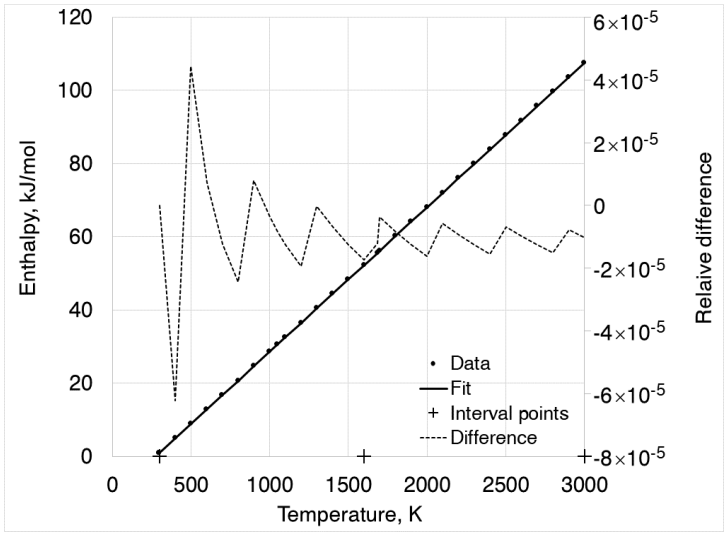


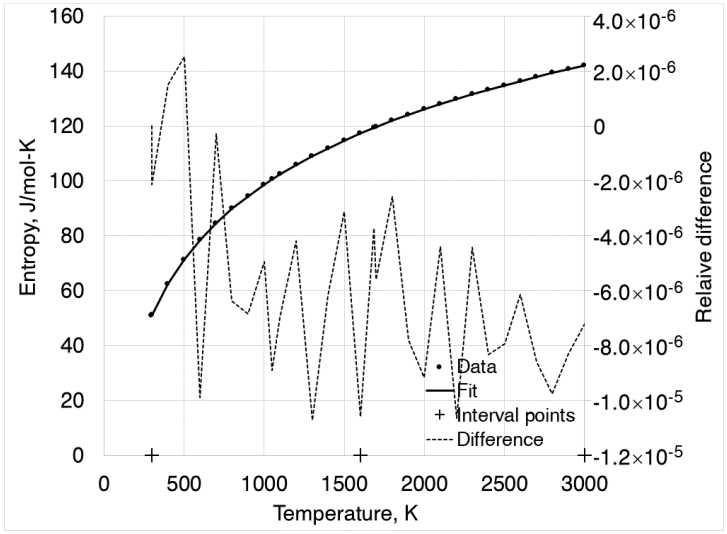


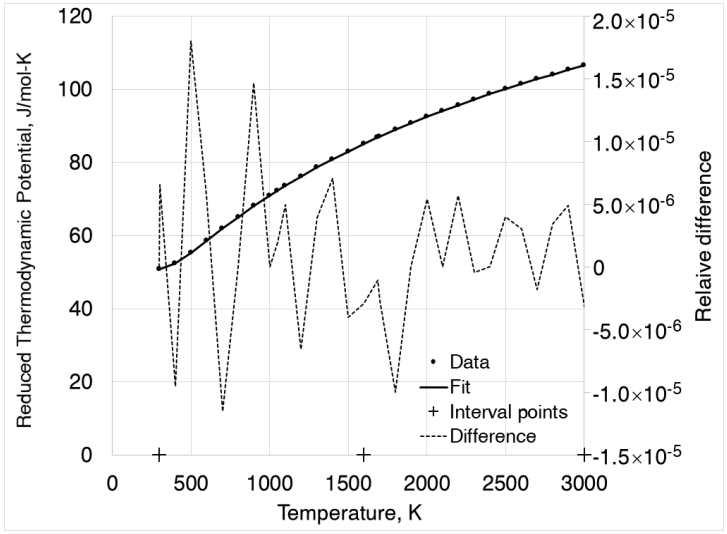
Strontium (Sr) – Sr1(l)

MW = 87.62 g/mol – $\Delta_f H^o$(298.15 K) = .911 kJ/mol – S^o(298.15 K) = 50.901 J/mol-K – p^o = 0.1 MPa



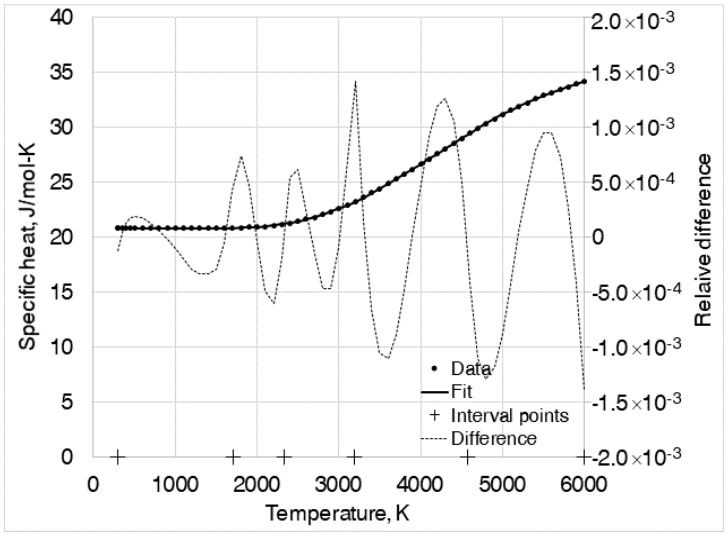


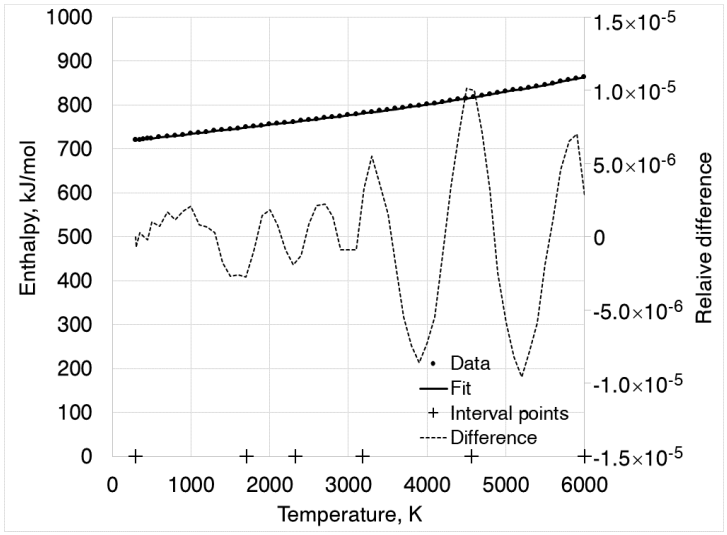


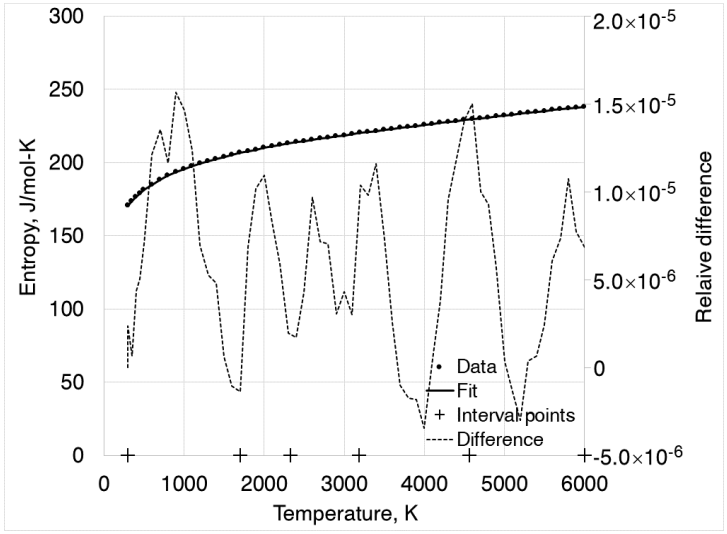


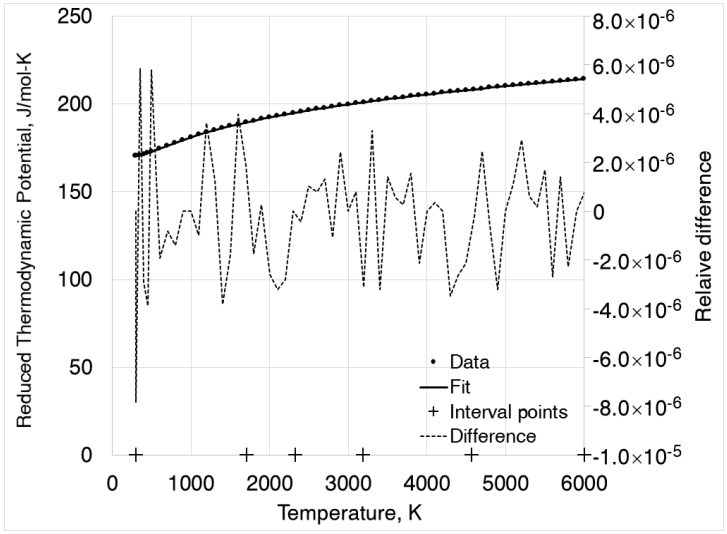
Strontium, Ion (Sr^+) – Sr1+(g)

MW = 87.62 g/mol – $\Delta_f H^o$(298.15 K) = 719.666 kJ/mol – S^o(298.15 K) = 170.403 J/mol-K – p^o = 0.1 MPa



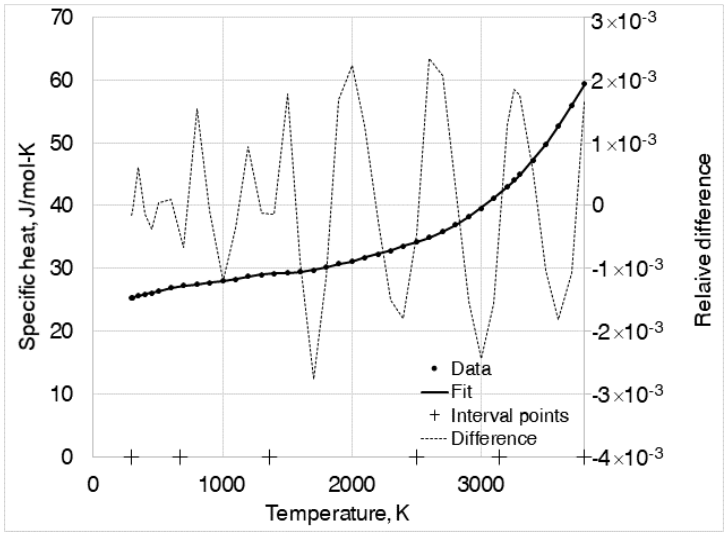


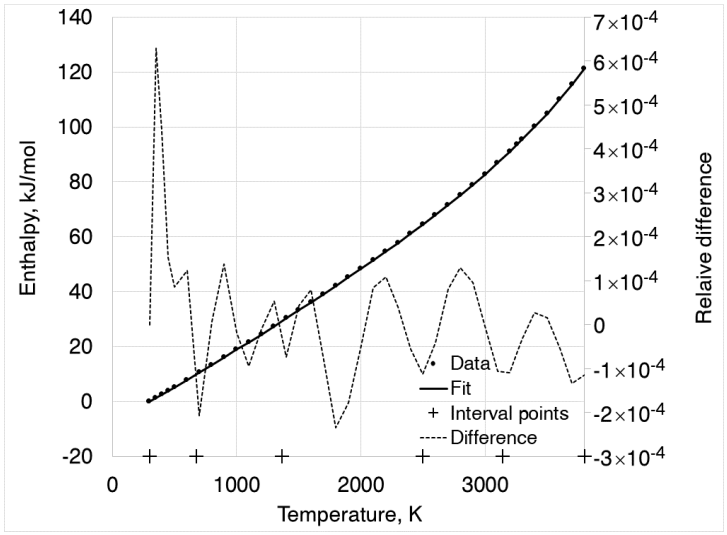


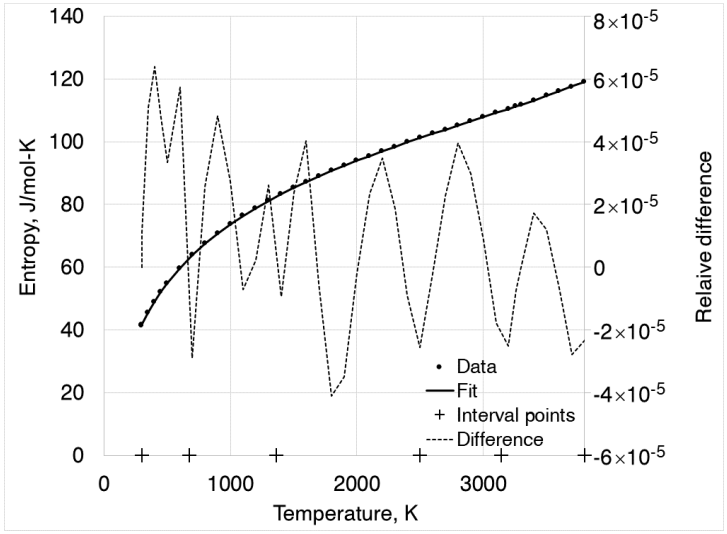


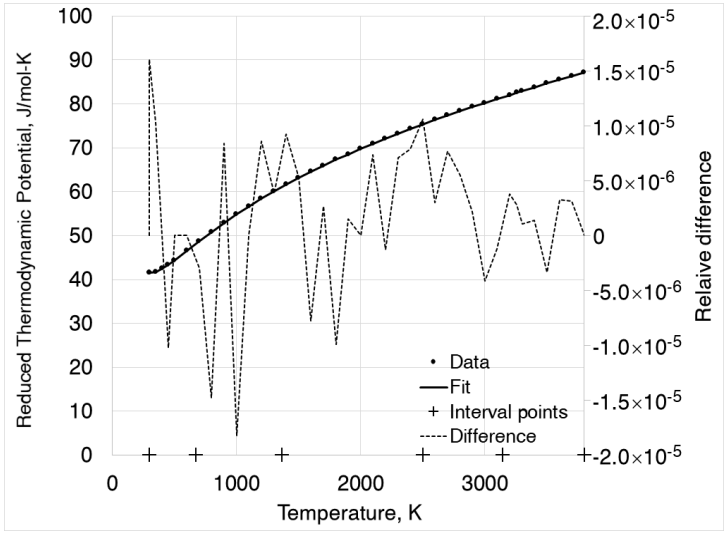
Tantalum (Ta) – Ta1(cr)

MW = 180.9479 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 41.471 J/mol-K – p^o = 0.1 MPa



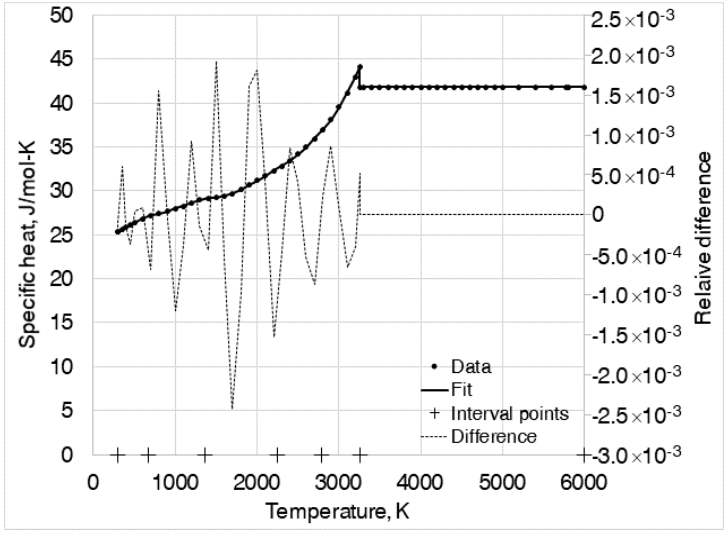


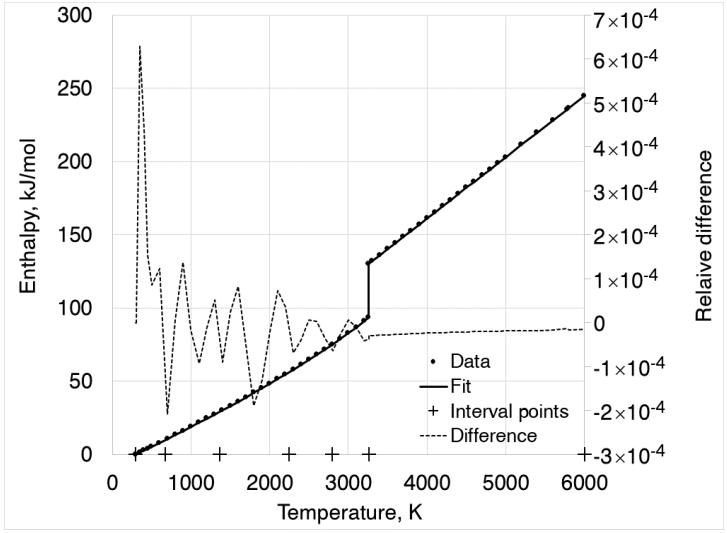


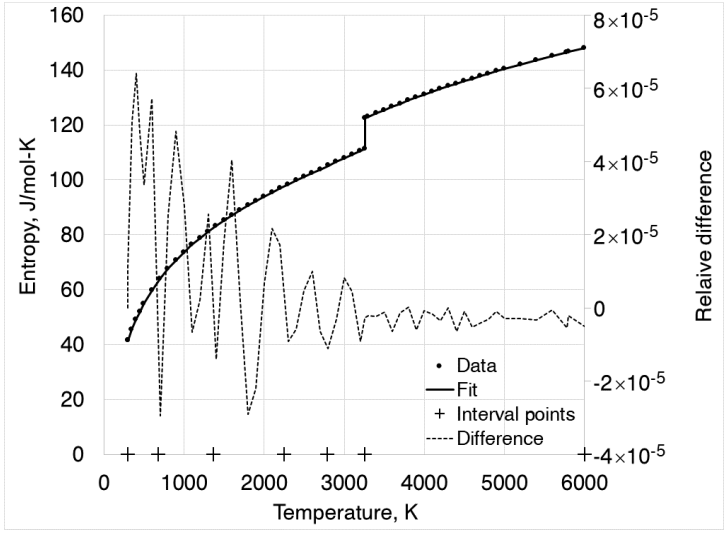


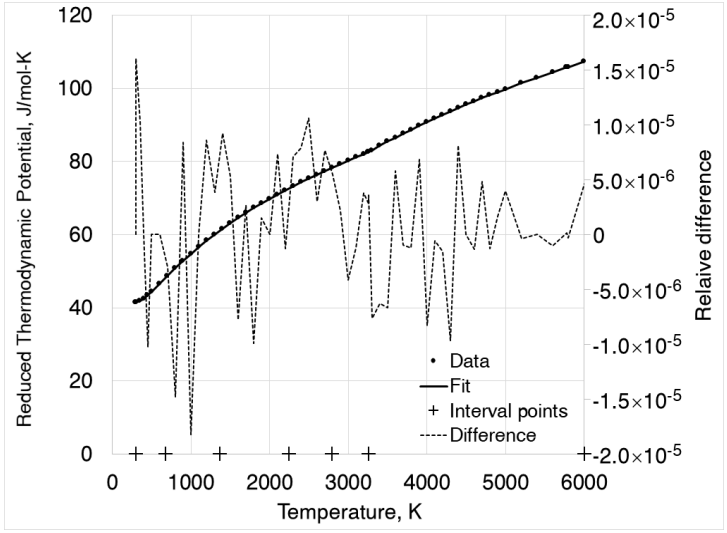
Tantalum (Ta) – Ta1(cr,l)

MW = 180.9479 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 41.471 J/mol-K – p^o = 0.1 MPa



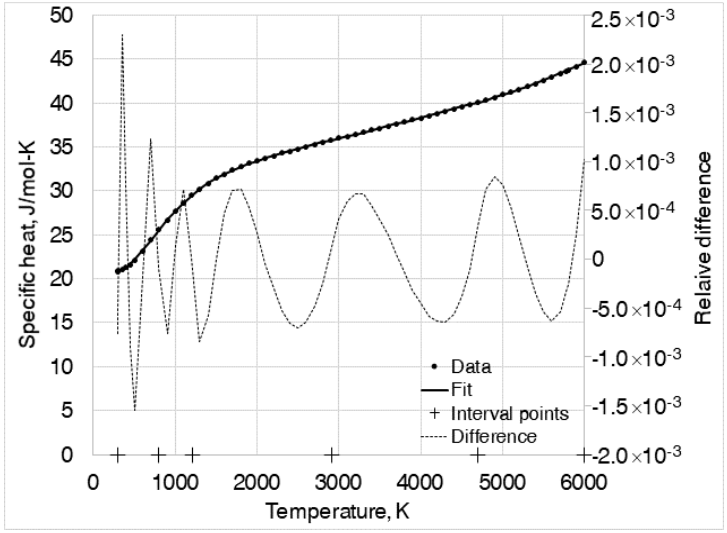


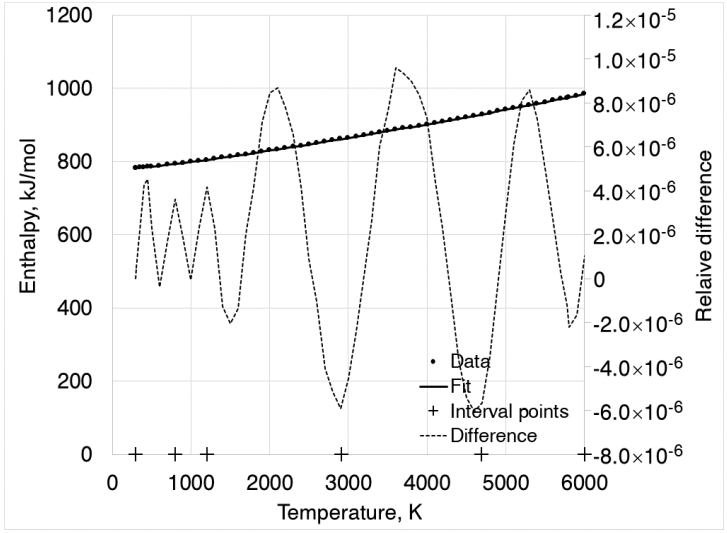


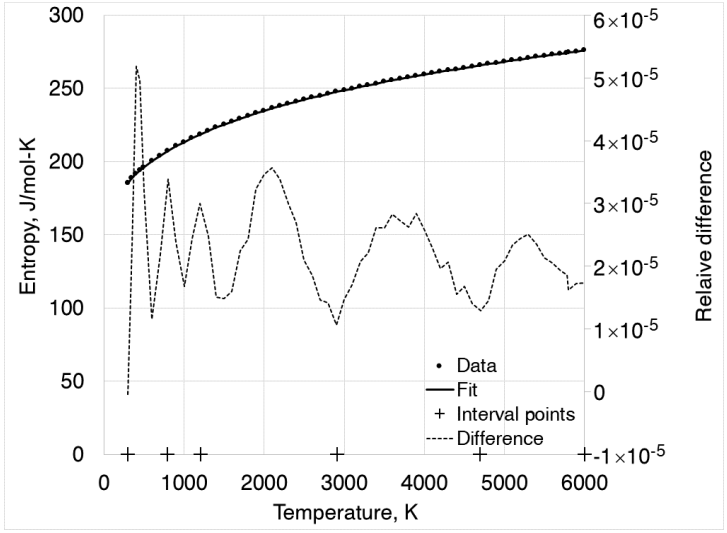


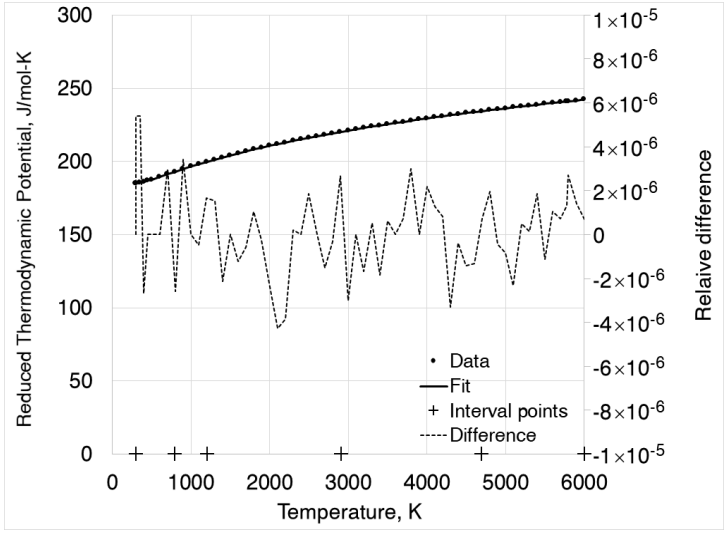
Tantalum (Ta) – Ta1(g)

MW = 180.9479 g/mol – $\Delta_f H^o$(298.15 K) = 781.99 kJ/mol – S^o(298.15 K) = 185.219 J/mol-K – p^o = 0.1 MPa



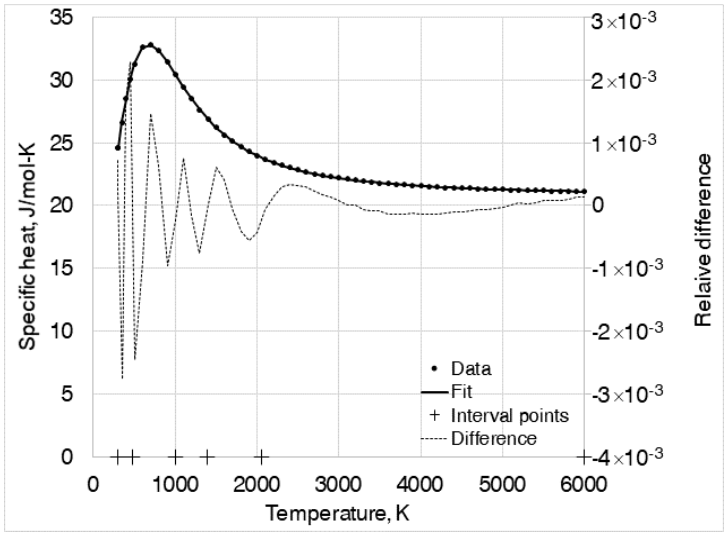


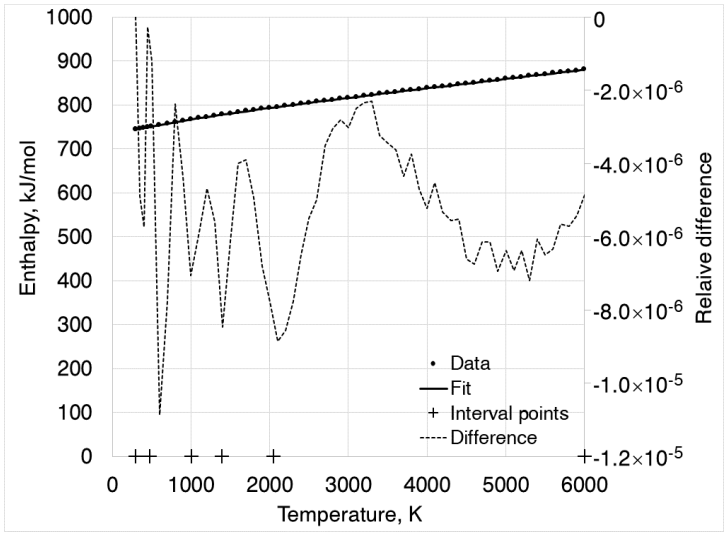


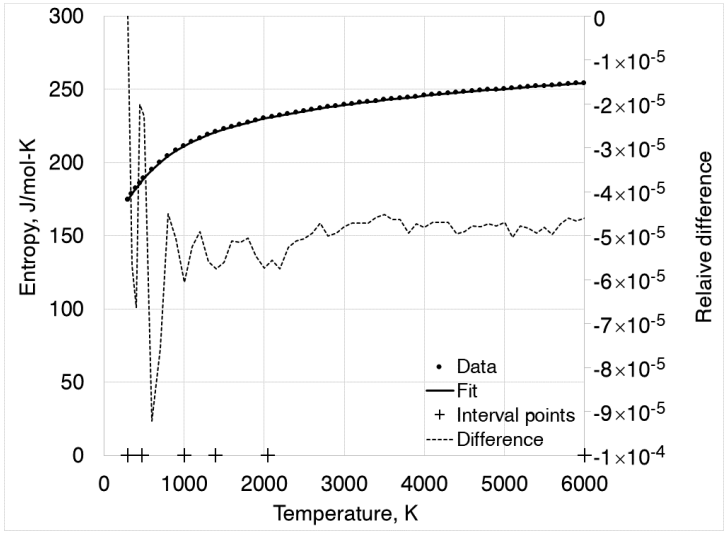


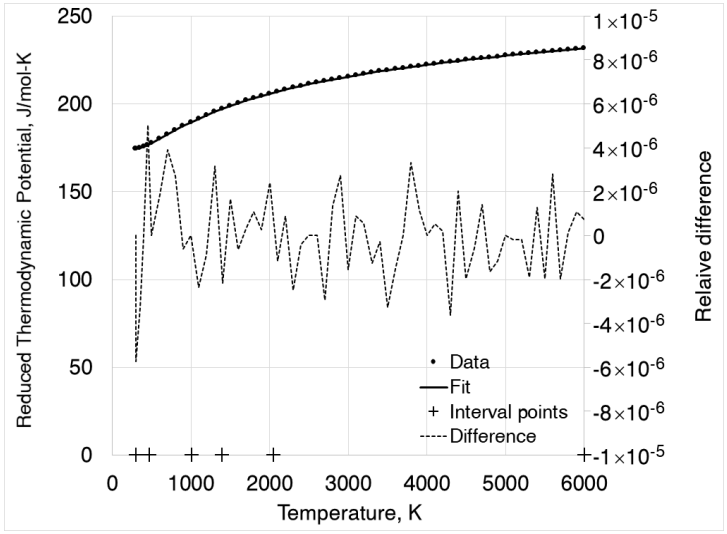
Tantalum, Ion (Ta^-) – Ta1-(g)

MW = 180.9484 g/mol – $\Delta_f H^o$(298.15 K) = 744.941 kJ/mol – S^o(298.15 K) = 174.56 J/mol-K – p^o = 0.1 MPa



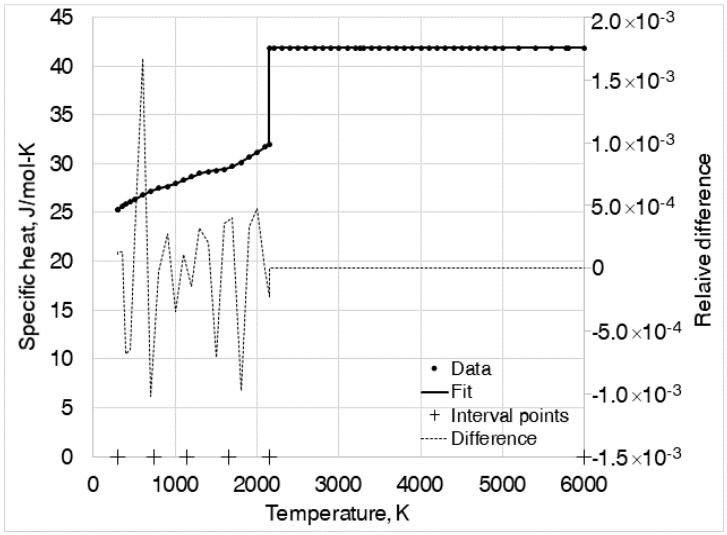


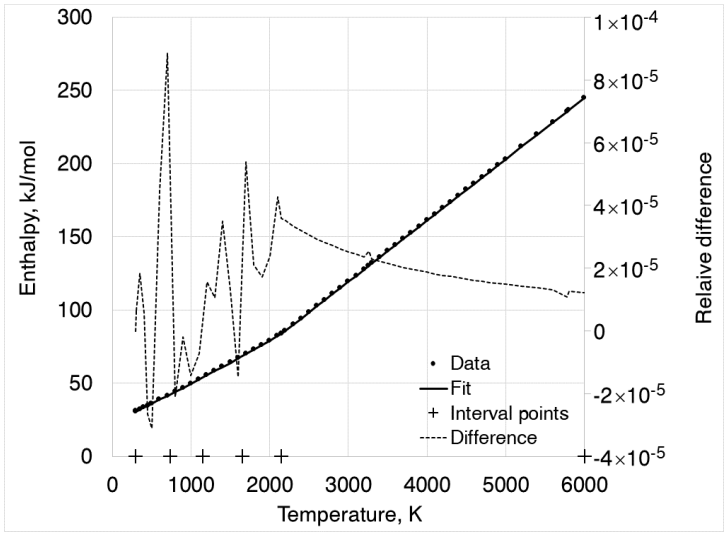


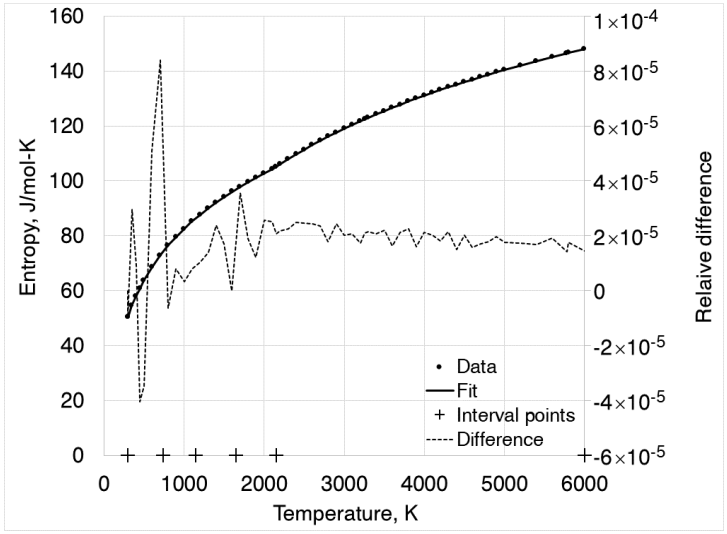


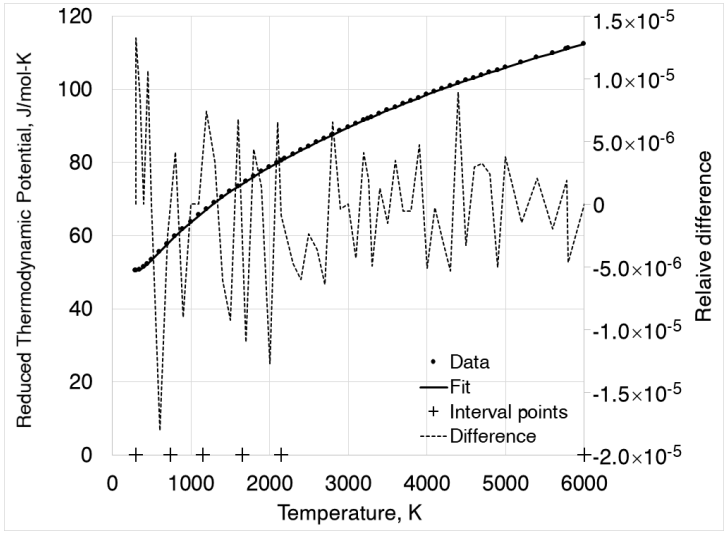
Tantalum (Ta) – Ta1(l)

MW = 180.9479 g/mol – $\Delta_f H^o$(298.15 K) = 30.795 kJ/mol – S^o(298.15 K) = 50.366 J/mol-K – p^o = 0.1 MPa



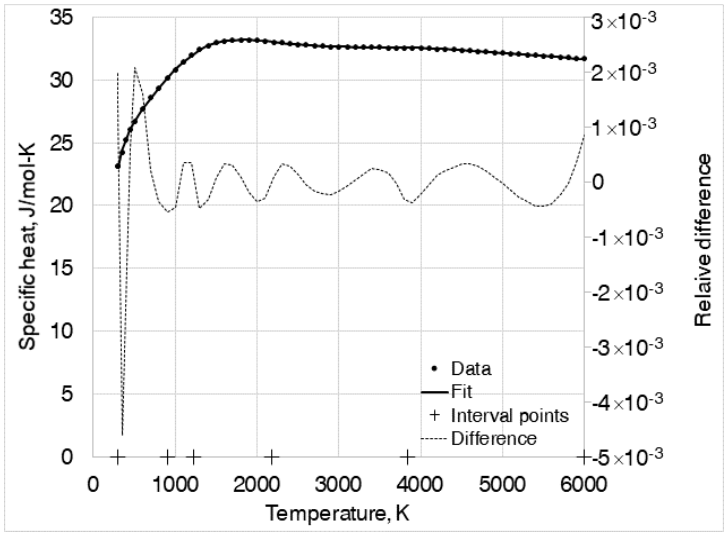


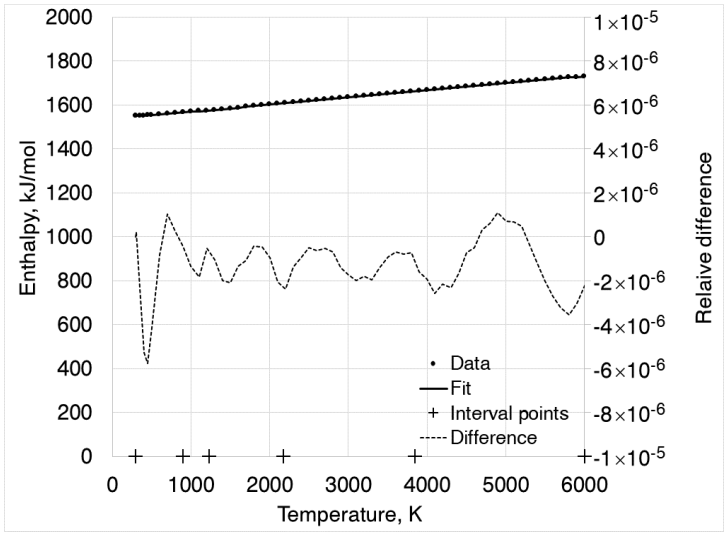


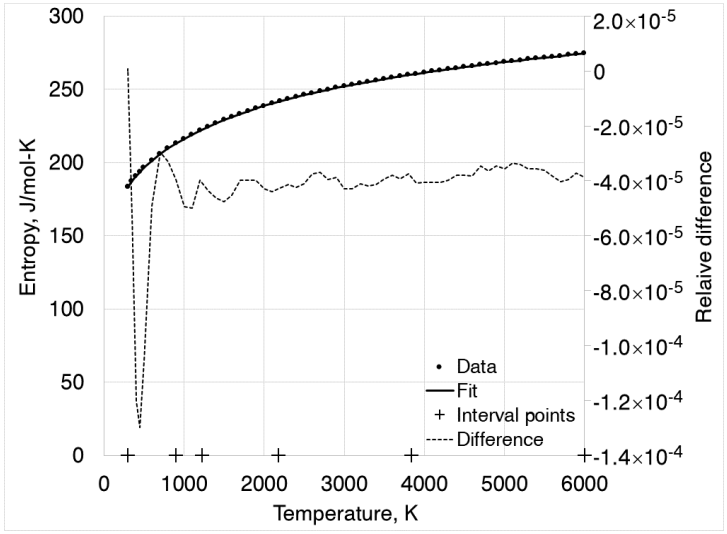


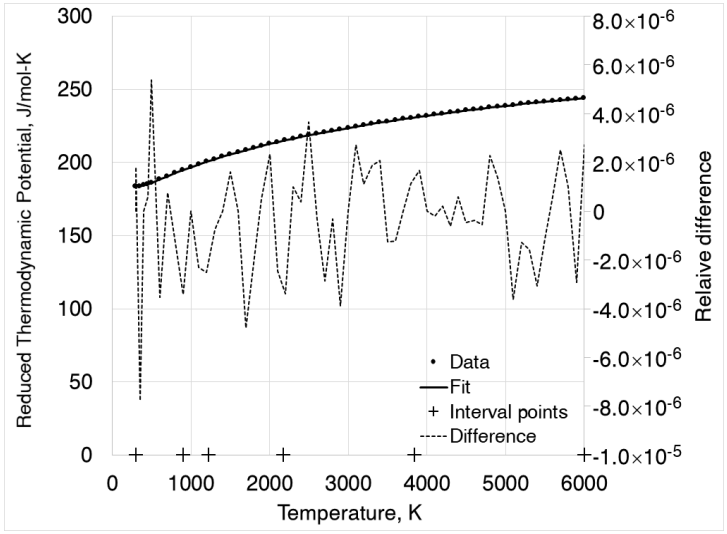
Tantalum, Ion (Ta^+) – Ta1+(g)

MW = 180.9474 g/mol – $\Delta_f H^o$(298.15 K) = 1549.15 kJ/mol – S^o(298.15 K) = 183.385 J/mol-K – p^o = 0.1 MPa



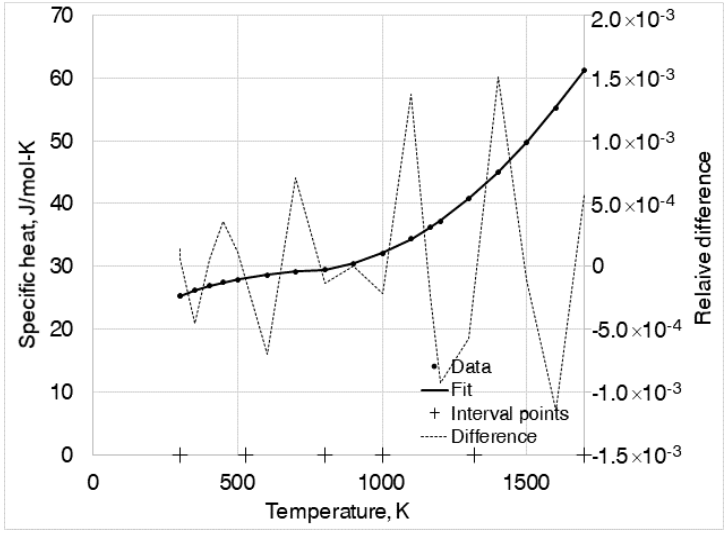


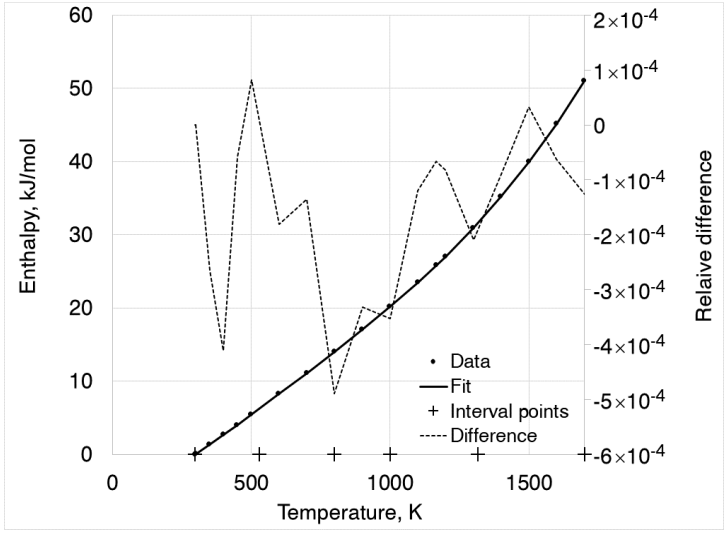


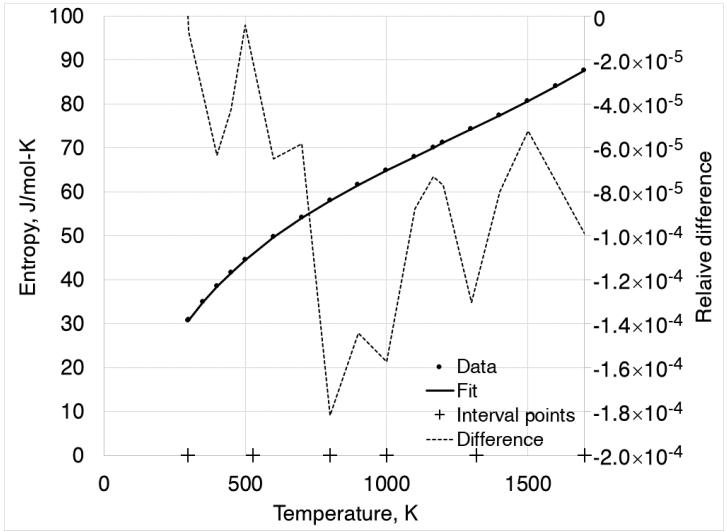


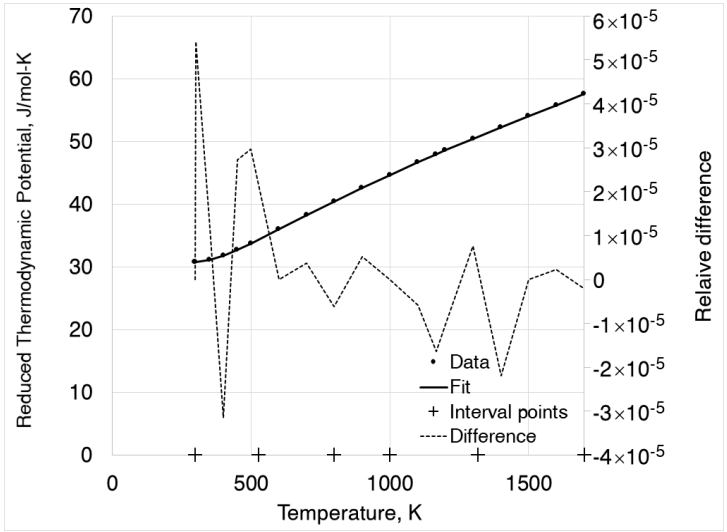
Titanium, Alpha (Ti) – Ti1(cr)_01

MW = 47.88 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 30.759 J/mol-K – p^o = 0.1 MPa



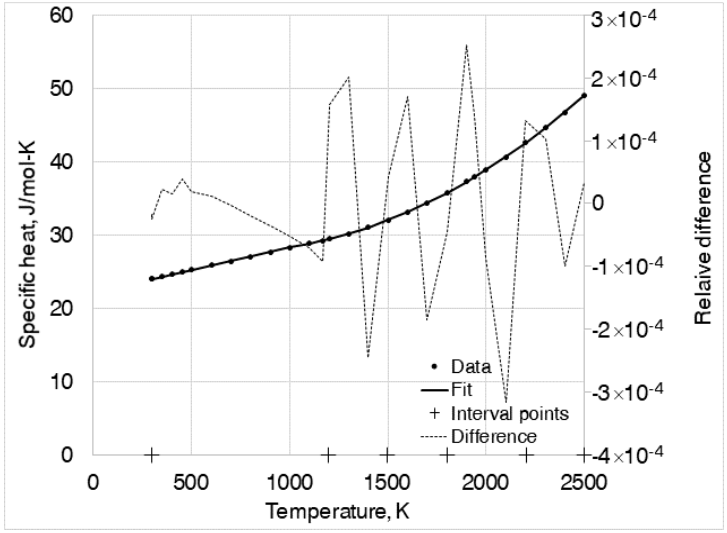


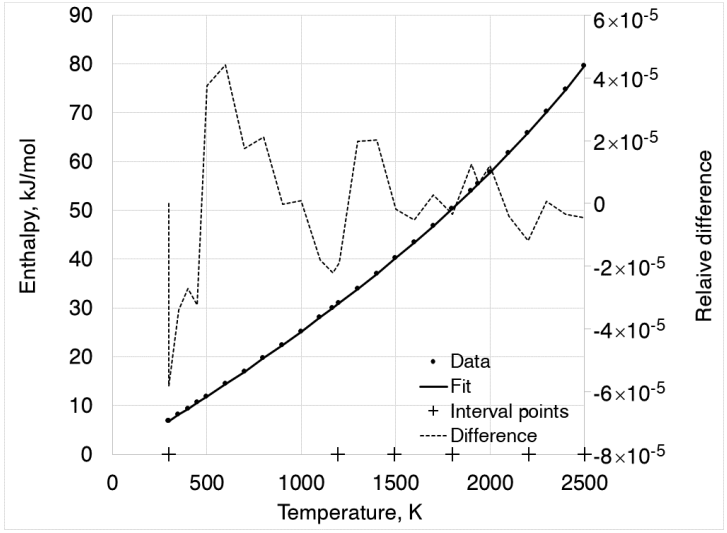


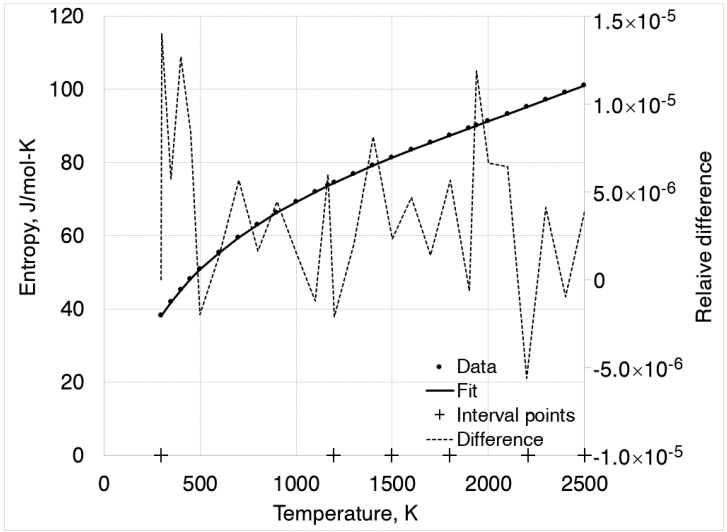


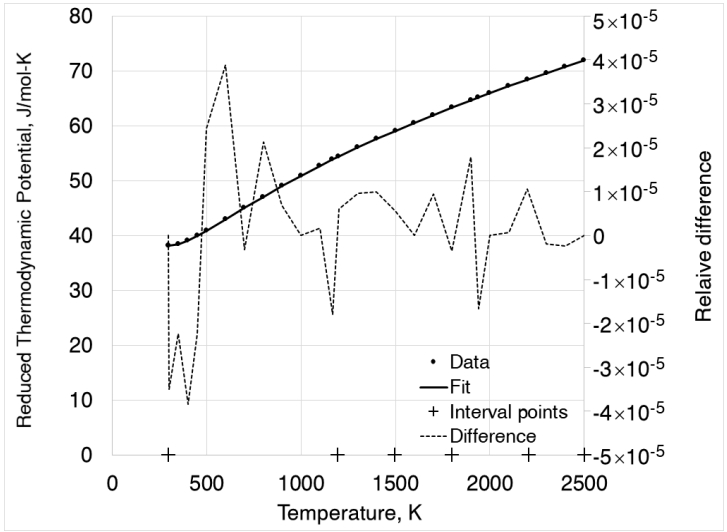
Titanium, Beta (Ti) – Ti1(cr)_02

MW = 47.88 g/mol – $\Delta_f H^o$(298.15 K) = 6.86 kJ/mol – S^o(298.15 K) = 38.122 J/mol-K – p^o = 0.1 MPa



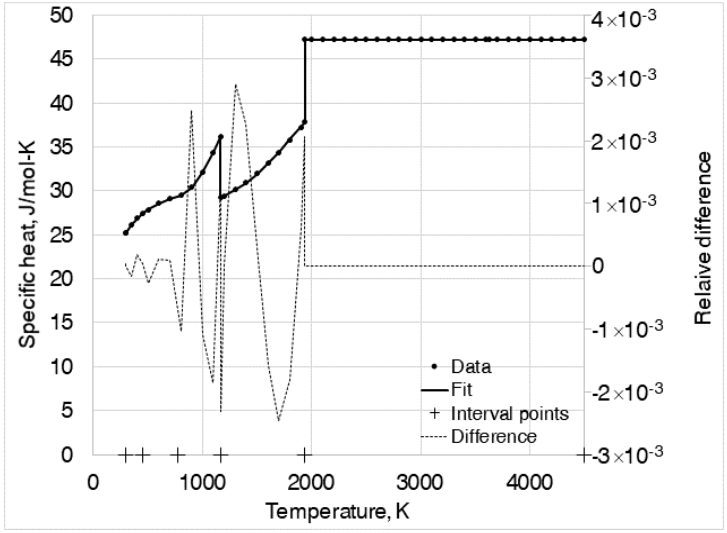


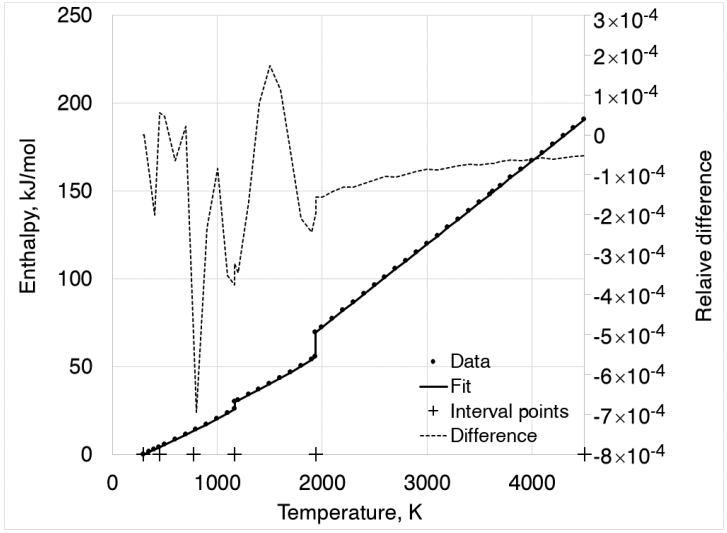


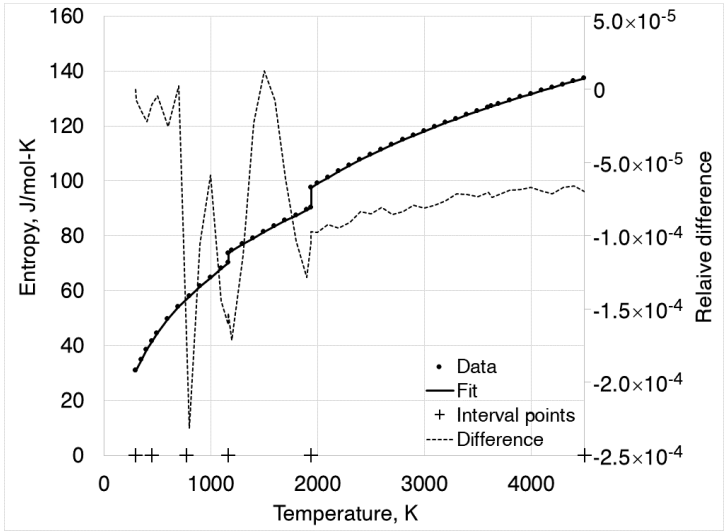


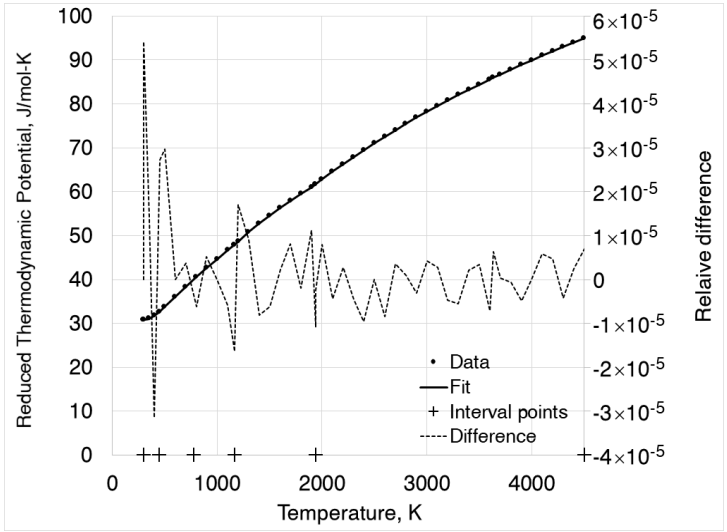
Titanium (Ti) – Ti1(cr,l)

MW = 47.88 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 30.759 J/mol-K – p^o = 0.1 MPa



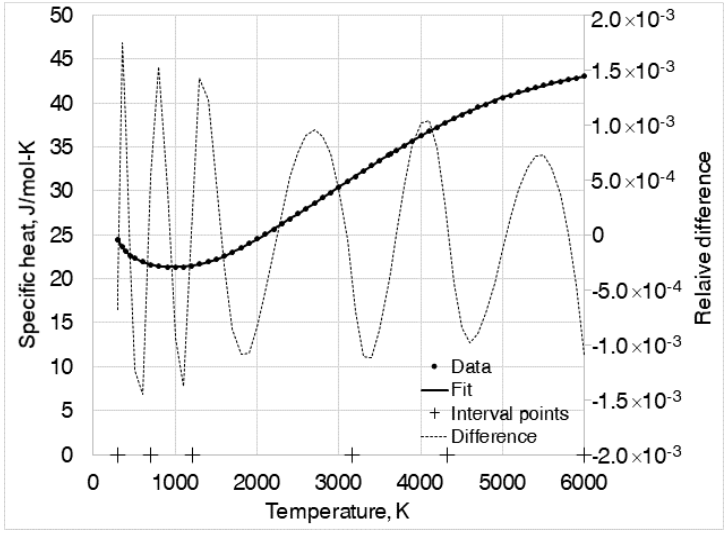


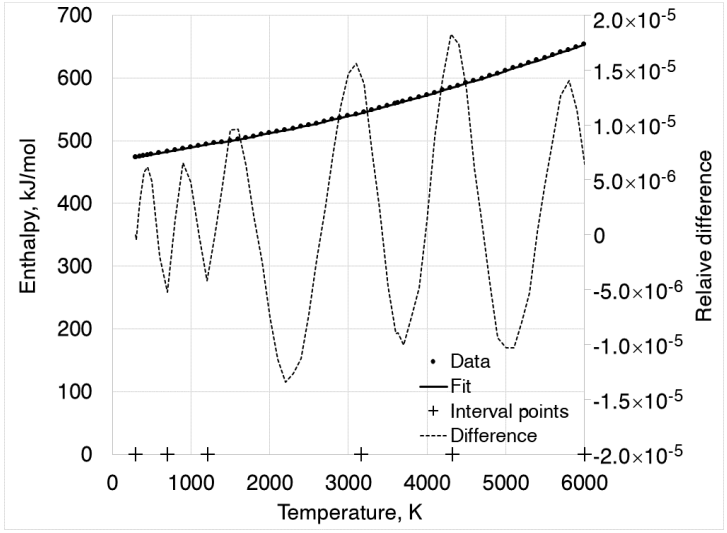


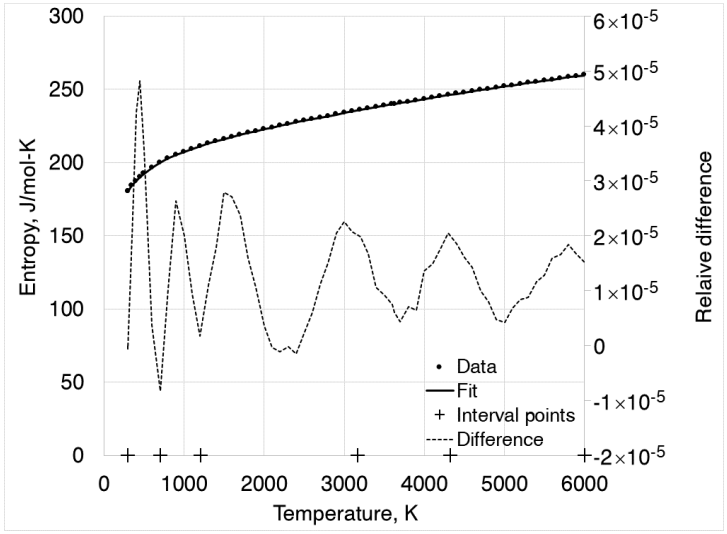


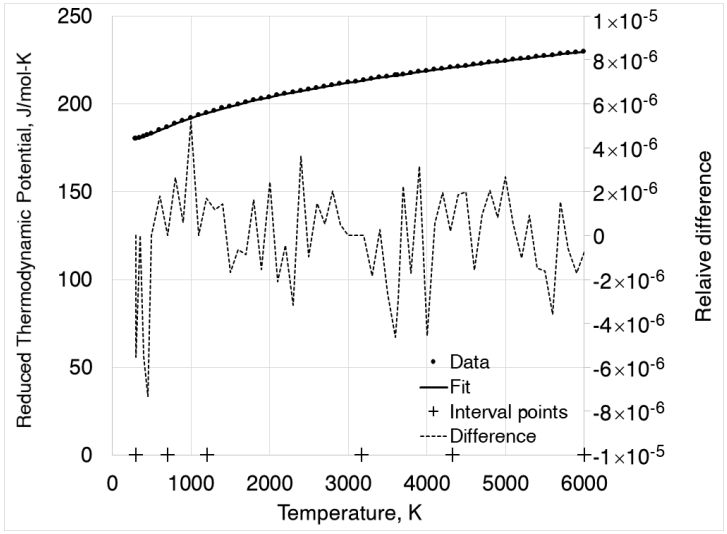
Titanium (Ti) – Ti1(g)

MW = 47.88 g/mol – $\Delta_f H^o$(298.15 K) = 473.629 kJ/mol – S^o(298.15 K) = 180.297 J/mol-K – p^o = 0.1 MPa



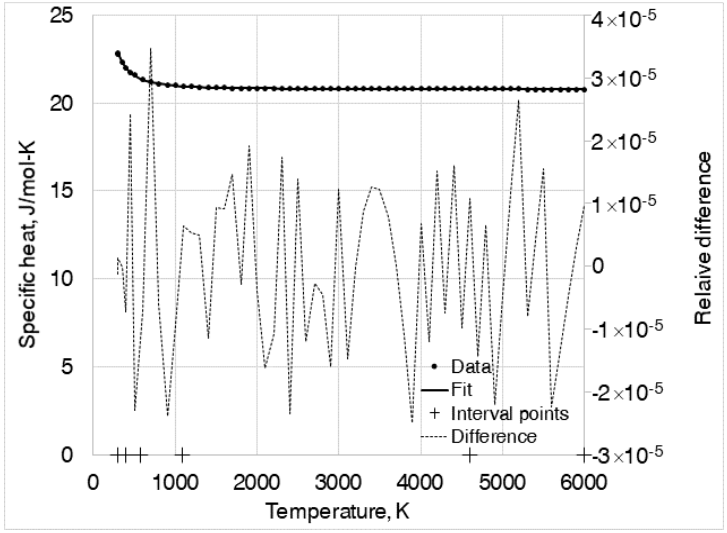


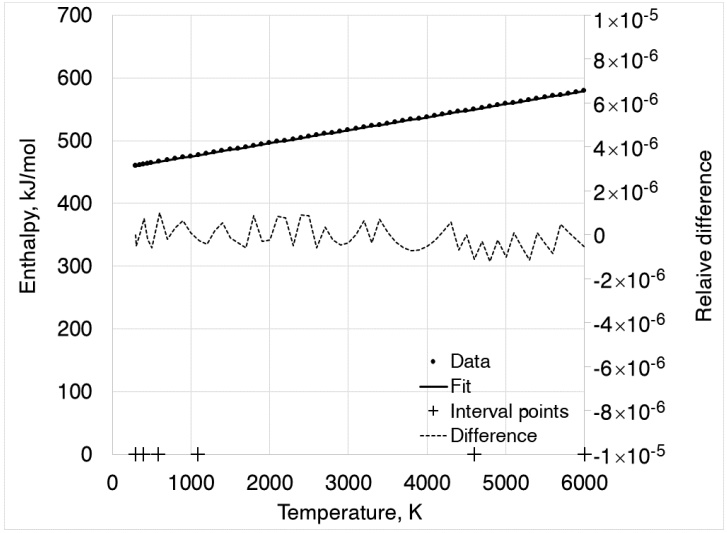


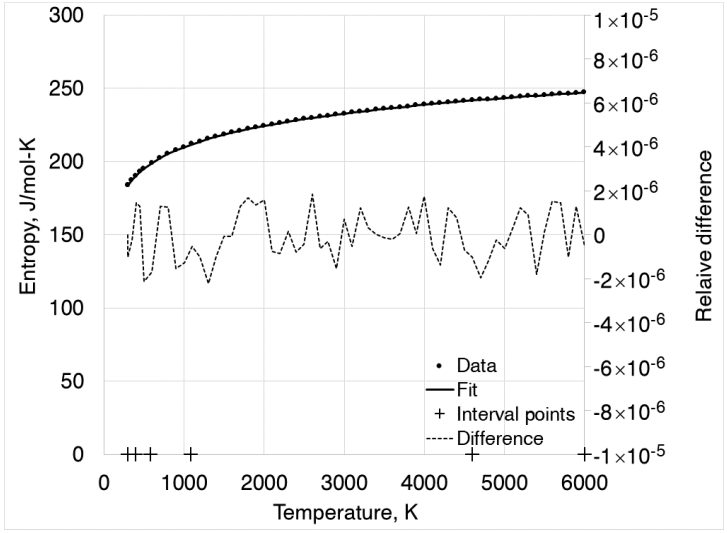


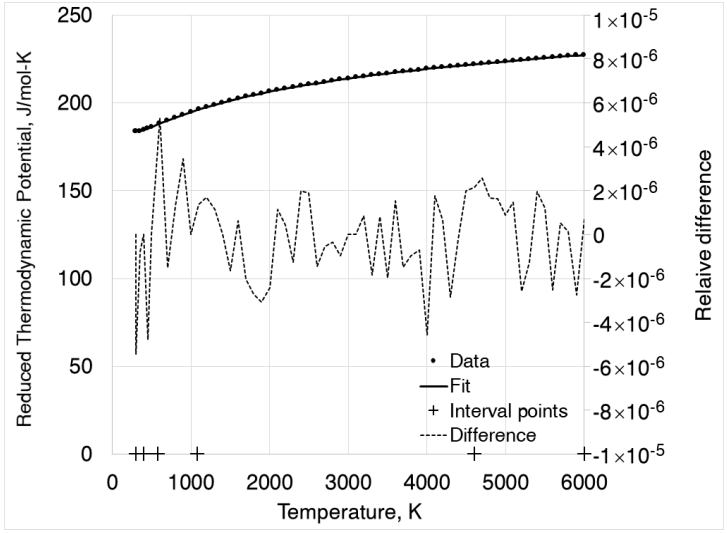
Titanium, Ion (Ti^-) – Ti1-(g)

MW = 47.88 g/mol – $\Delta_f H^o$(298.15 K) = 459.833 kJ/mol – S^o(298.15 K) = 183.715 J/mol-K – p^o = 0.1 MPa



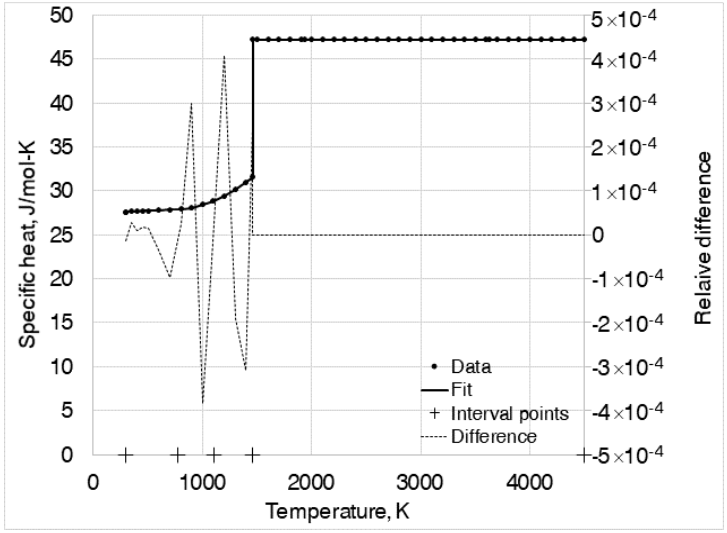


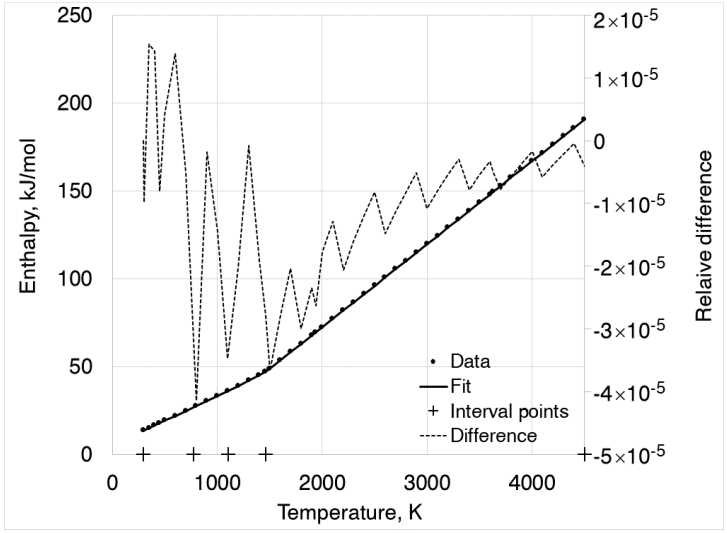


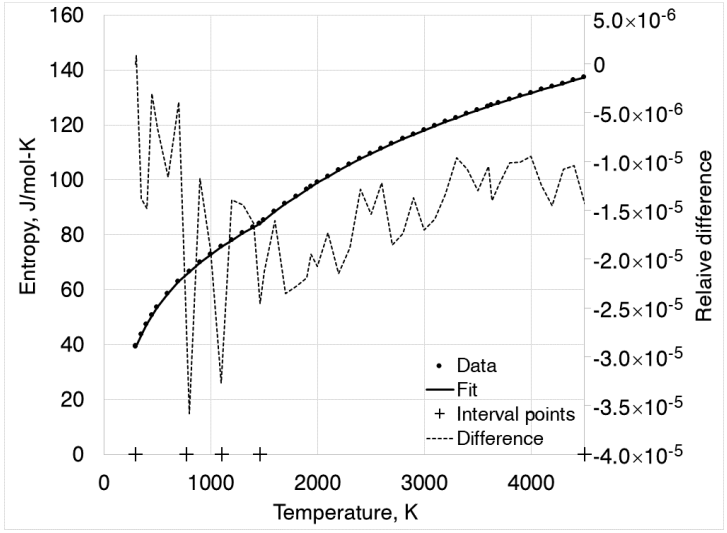


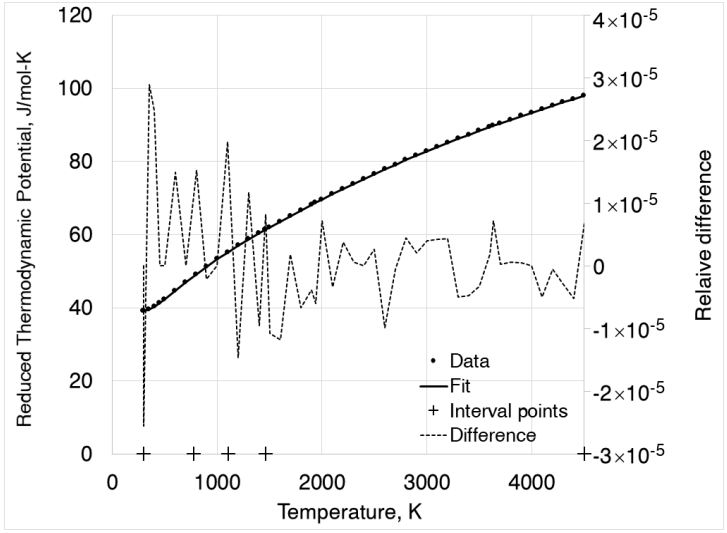
Titanium (Ti) – Ti1(l)

MW = 47.88 g/mol – $\Delta_f H^o$(298.15 K) = 13.652 kJ/mol – S^o(298.15 K) = 39.181 J/mol-K – p^o = 0.1 MPa



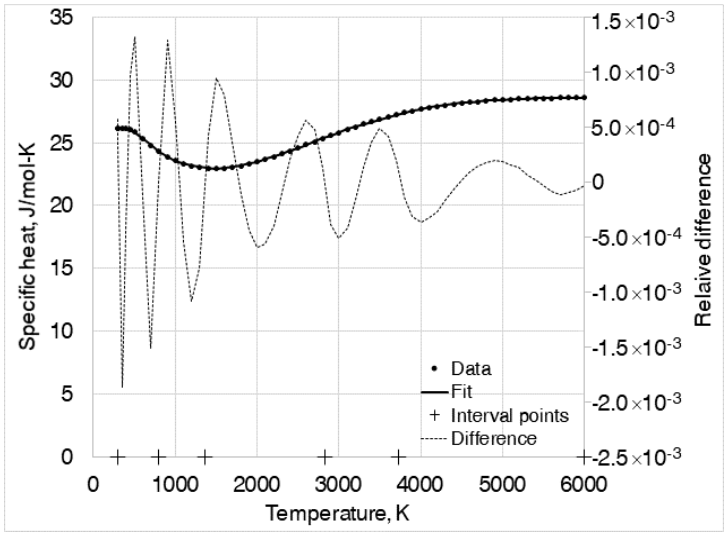


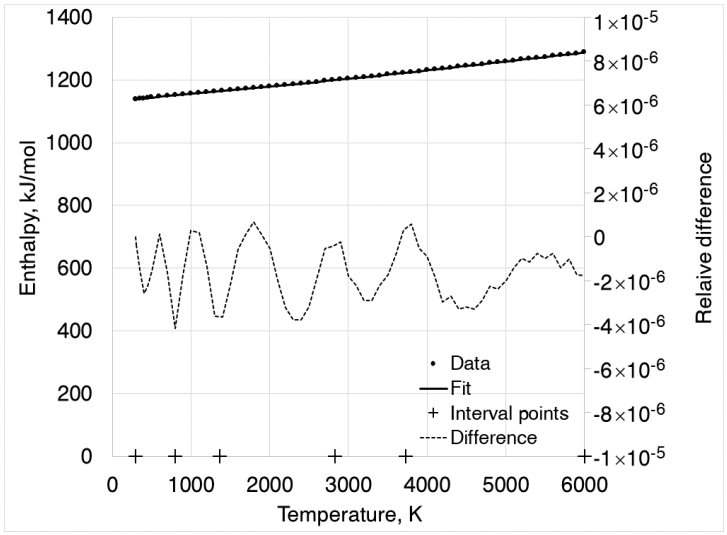


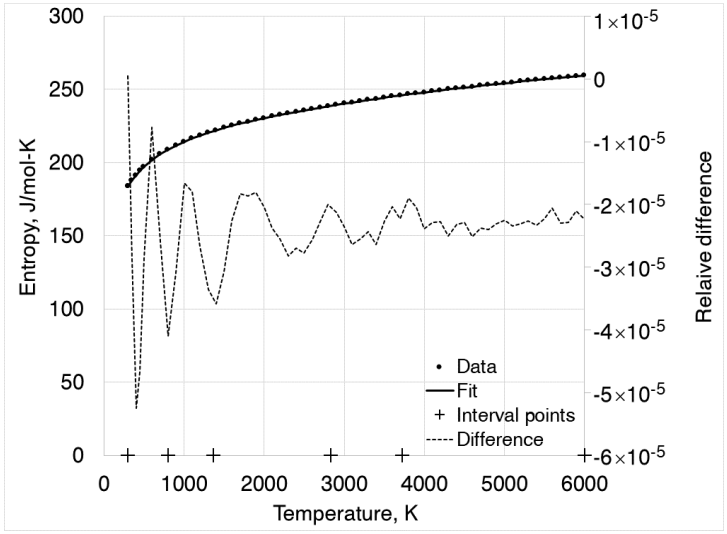


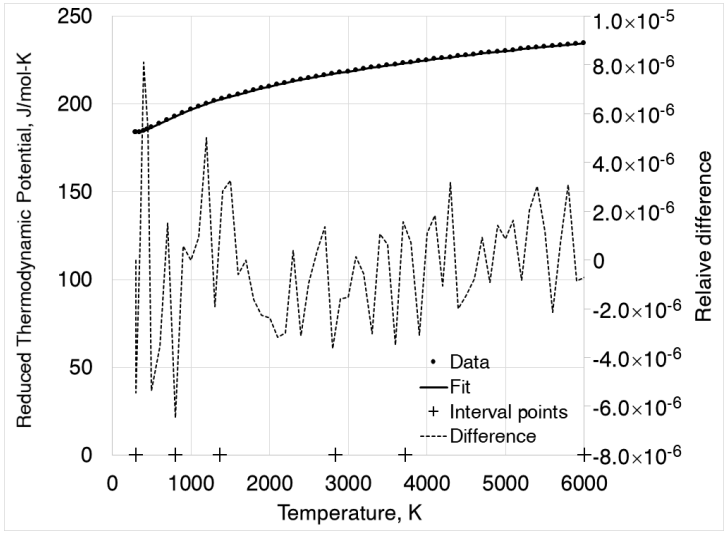
Titanium, Ion (Ti^+) – Ti1+(g)

MW = 47.88 g/mol – $\Delta_f H^o$(298.15 K) = 1138.252 kJ/mol – S^o(298.15 K) = 183.594 J/mol-K – p^o = 0.1 MPa



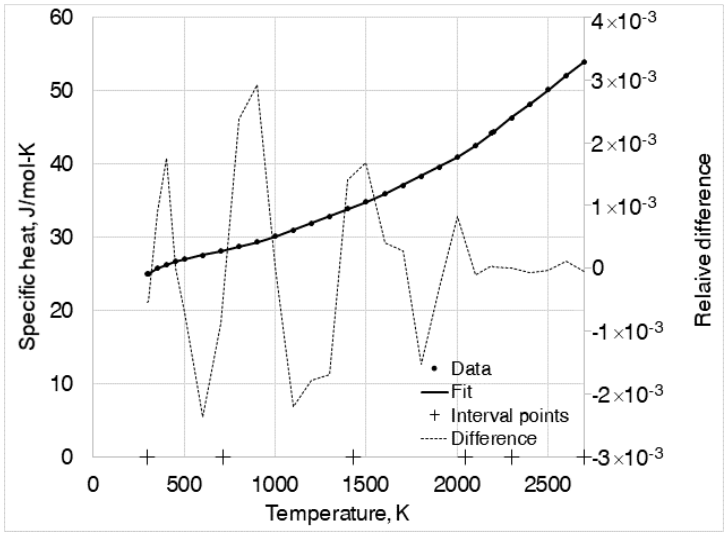


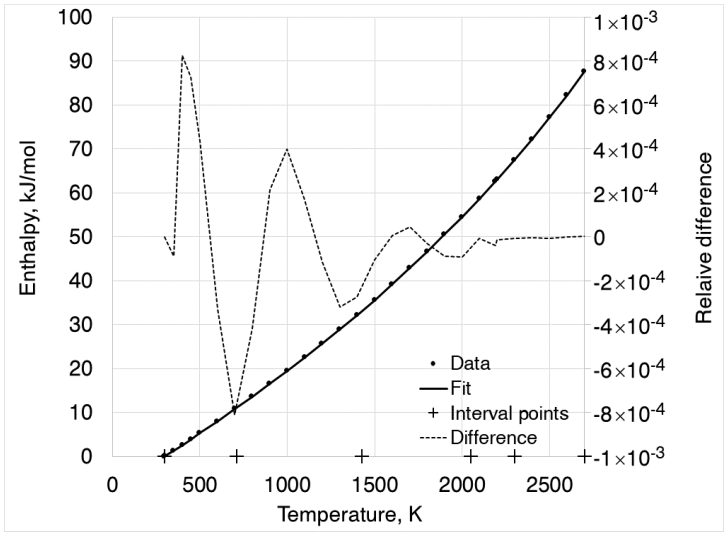


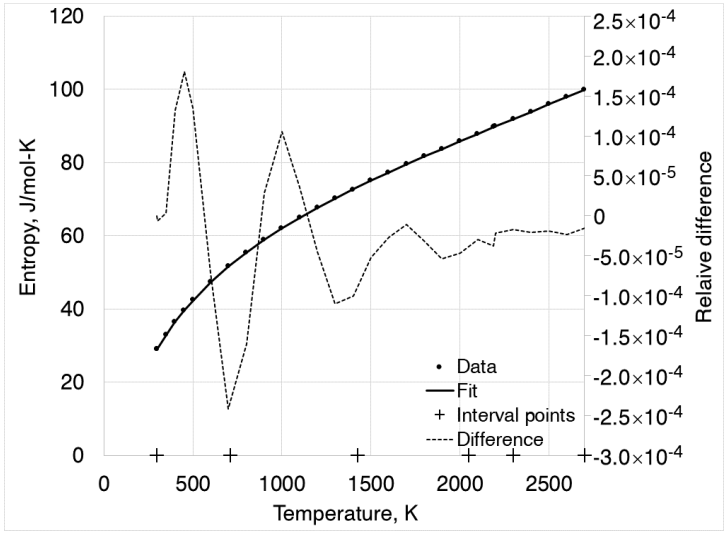


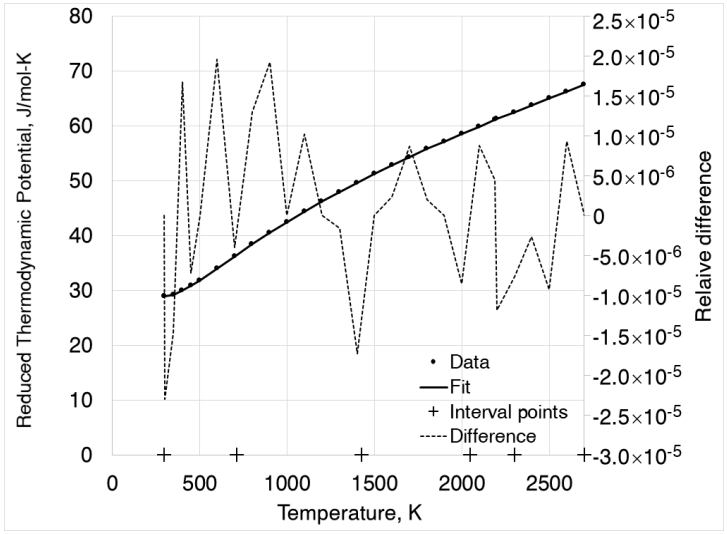
Vanadium (V) – V1(cr)

MW = 50.9415 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 28.936 J/mol-K – p^o = 0.1 MPa



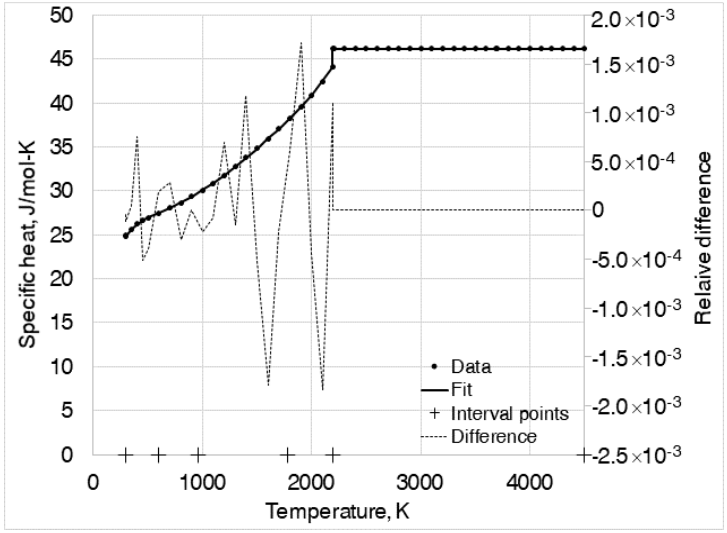


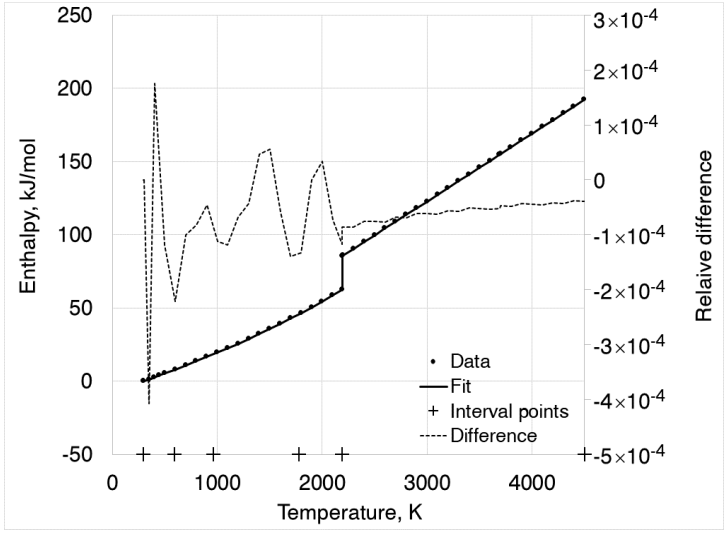


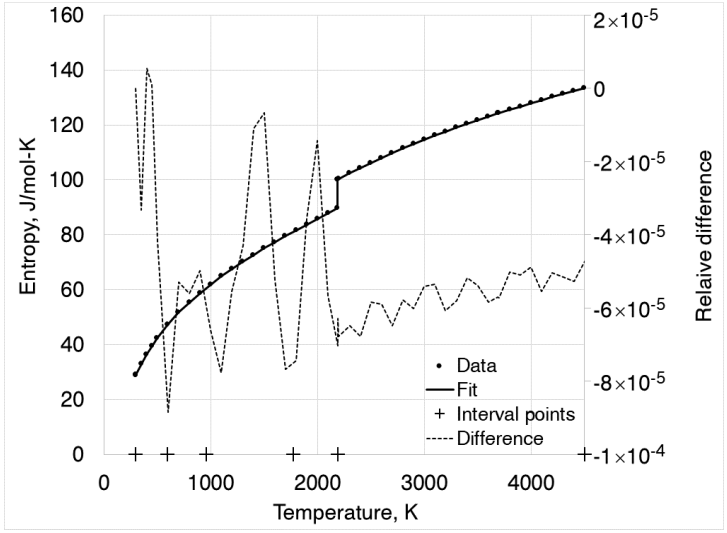


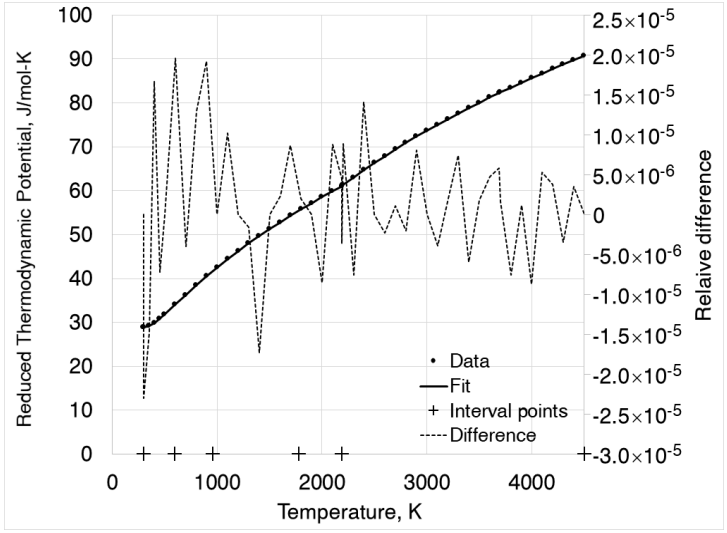
Vanadium (V) – V1(cr,l)

MW = 50.9415 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 28.936 J/mol-K – p^o = 0.1 MPa



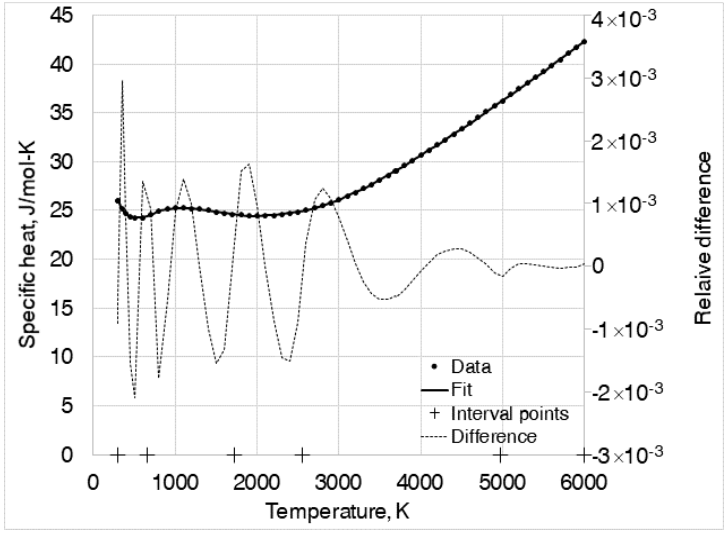


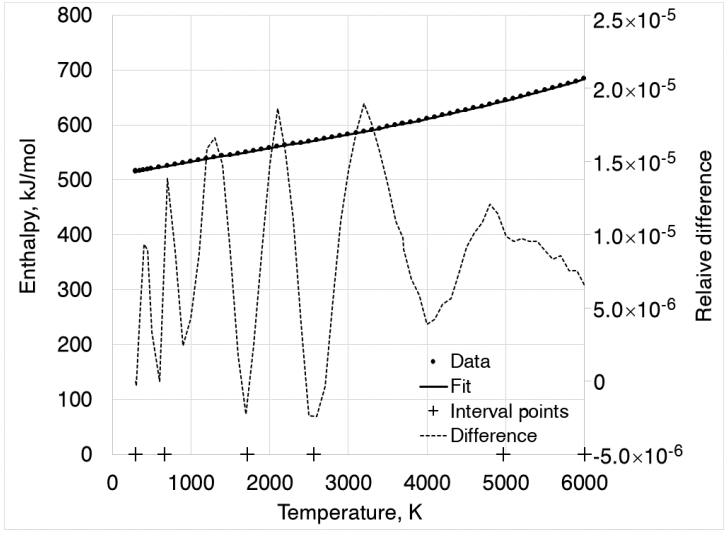


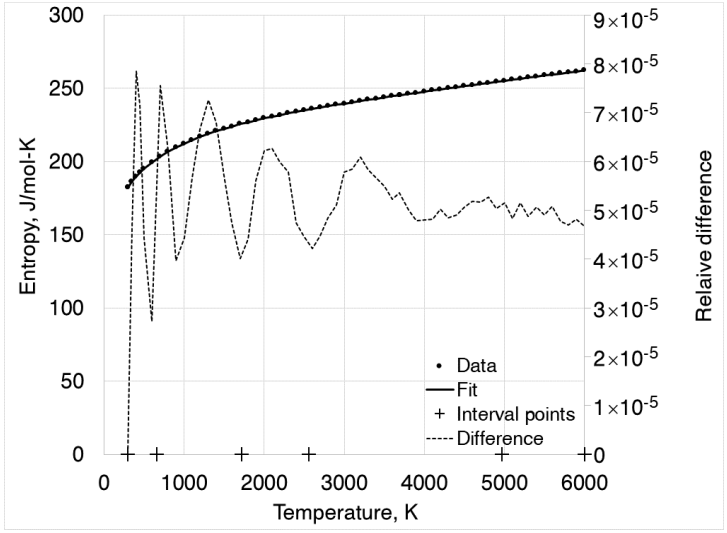


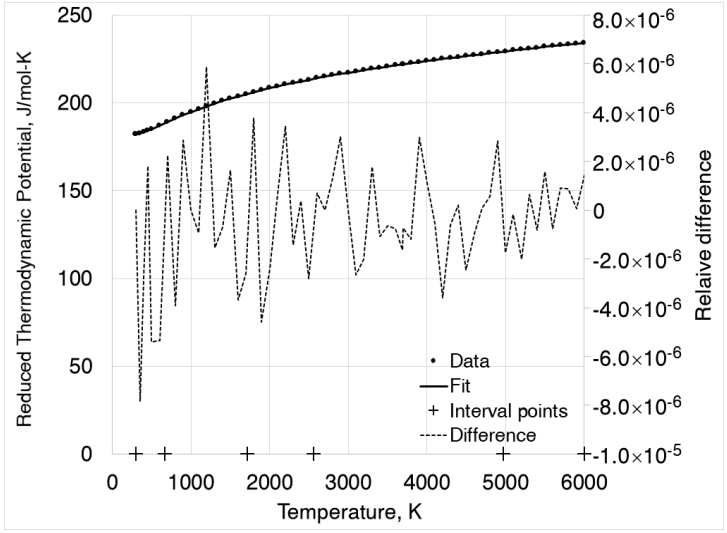
Vanadium (V) – V1(g)

MW = 50.9415 g/mol – $\Delta_f H^o$(298.15 K) = 515.469 kJ/mol – S^o(298.15 K) = 182.298 J/mol-K – p^o = 0.1 MPa



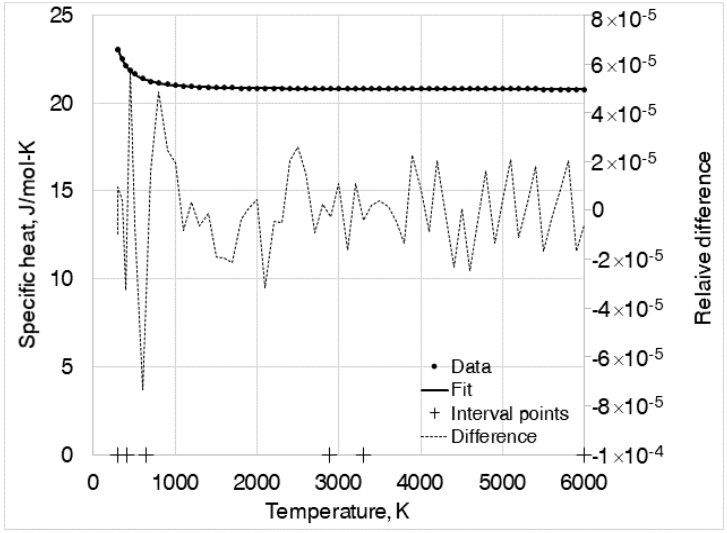


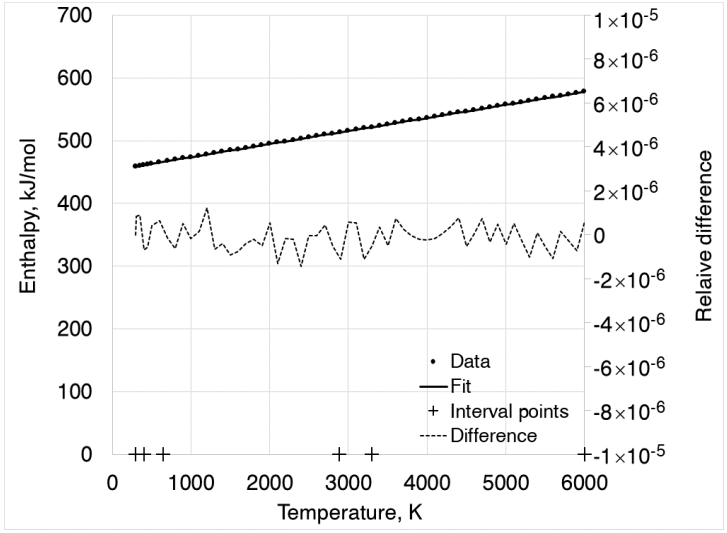


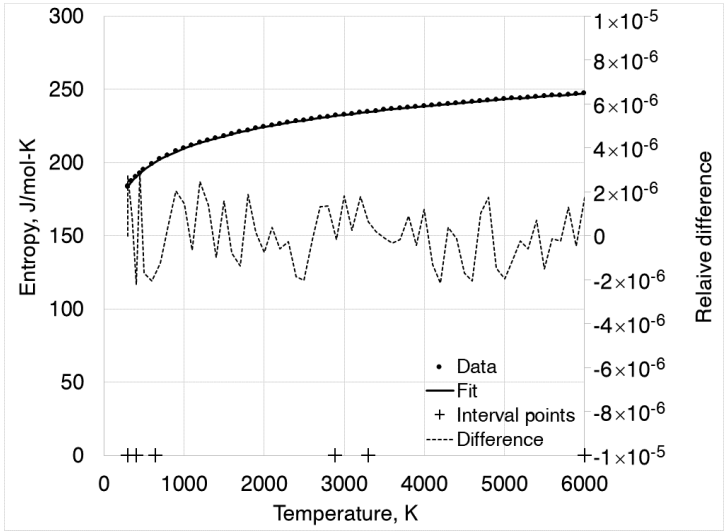


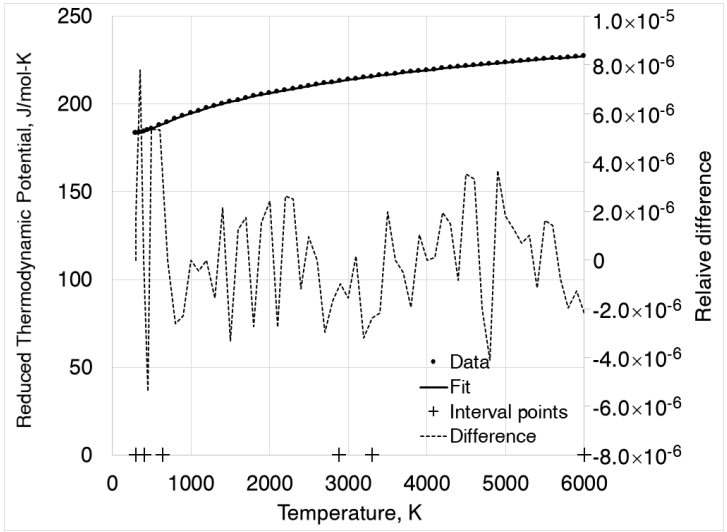
Vanadium, Ion (V^-) – V1-(g)

MW = 50.942 g/mol – $\Delta_f H^o$(298.15 K) = 458.588 kJ/mol – S^o(298.15 K) = 183.436 J/mol-K – p^o = 0.1 MPa



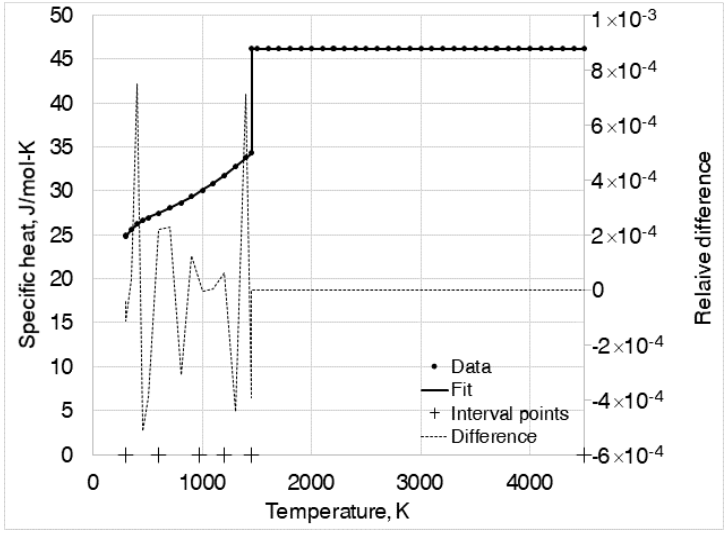


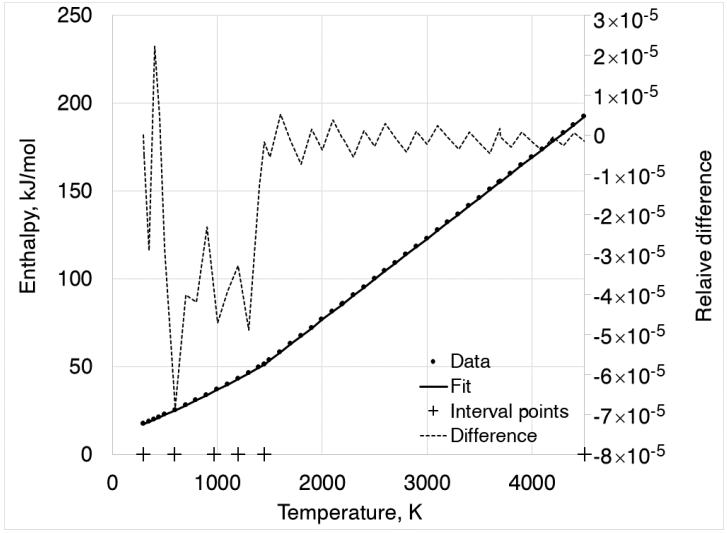


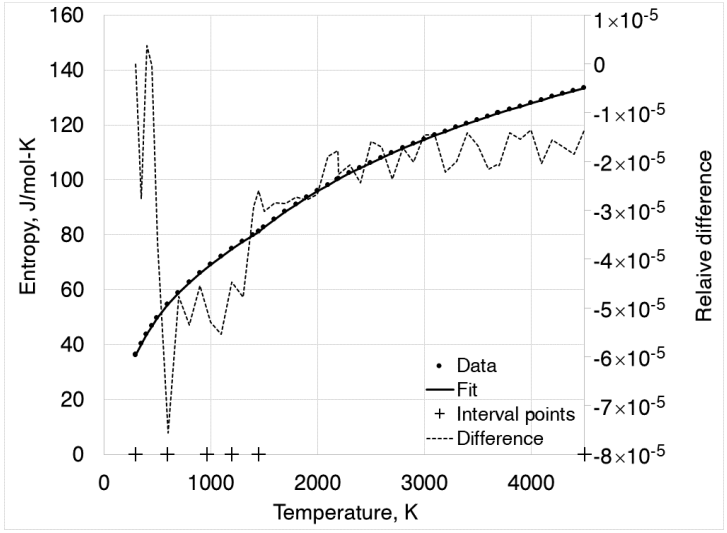


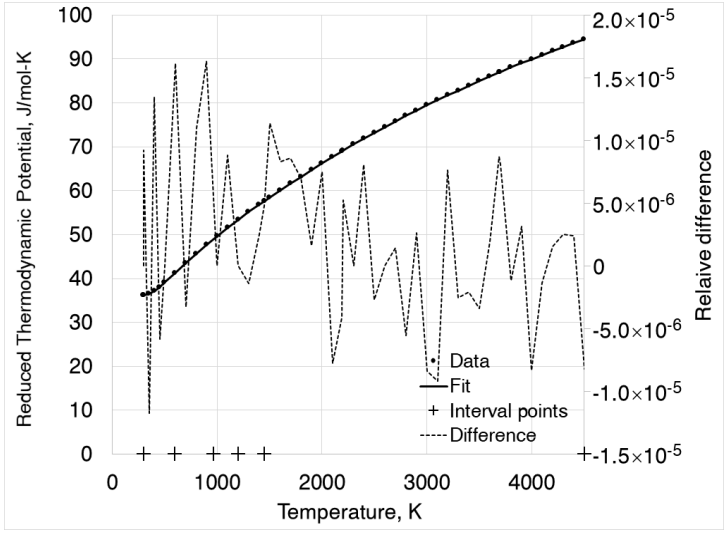
Vanadium (V) – V1(l)

MW = 50.9415 g/mol – $\Delta_f H^o$(298.15 K) = 17.288 kJ/mol – S^o(298.15 K) = 36.136 J/mol-K – p^o = 0.1 MPa



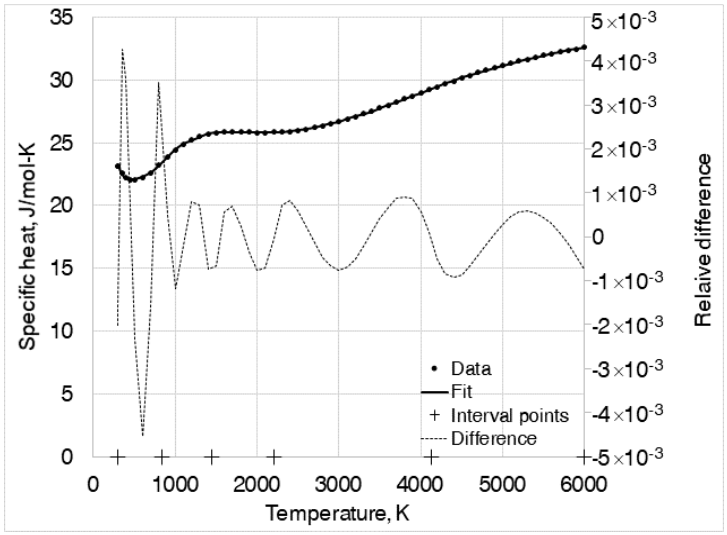


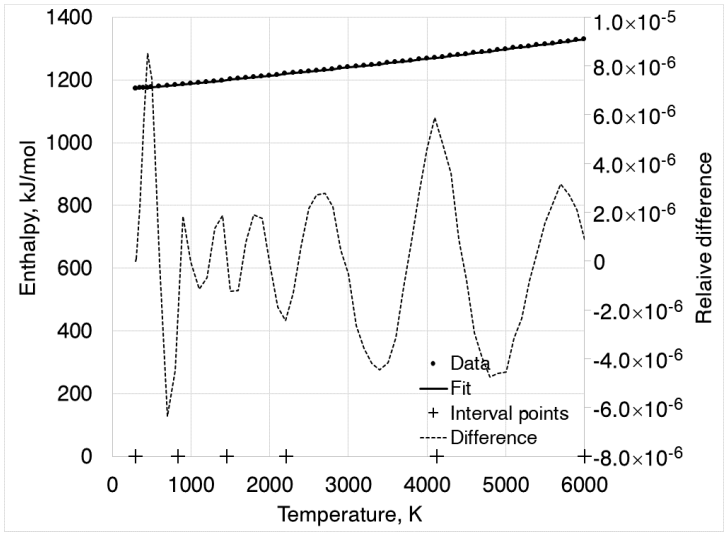


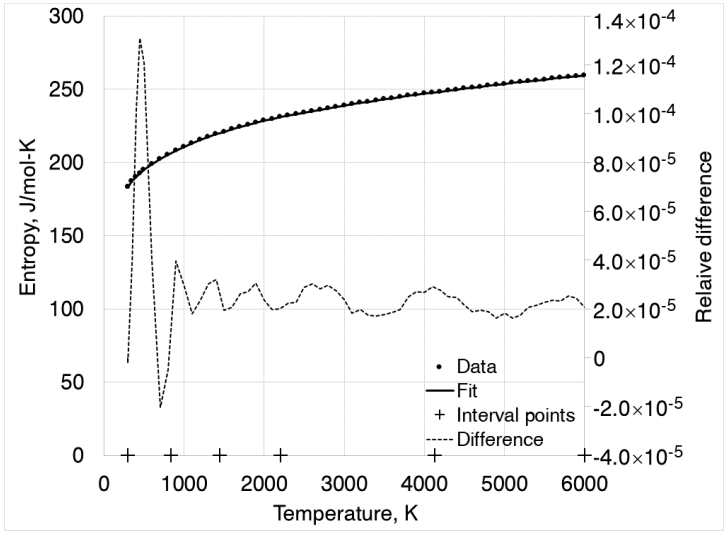


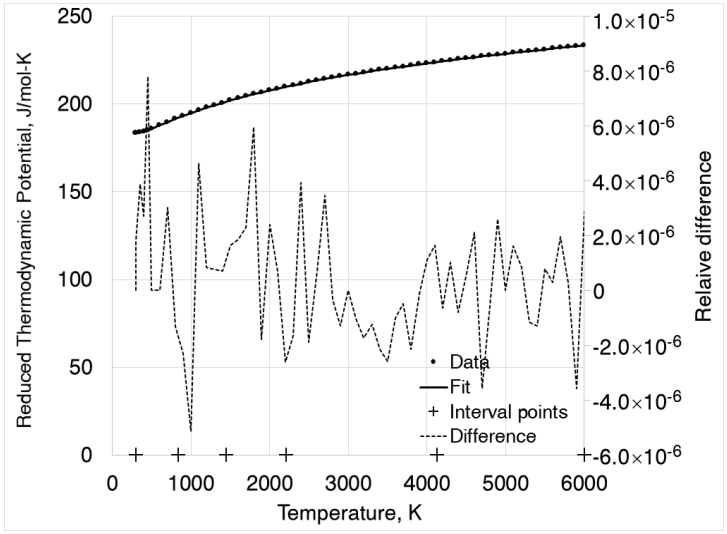
Vanadium, Ion (V^+) – V1+(g)

MW = 50.941 g/mol – $\Delta_f H^o$(298.15 K) = 1171.947 kJ/mol – S^o(298.15 K) = 183.376 J/mol-K – p^o = 0.1 MPa



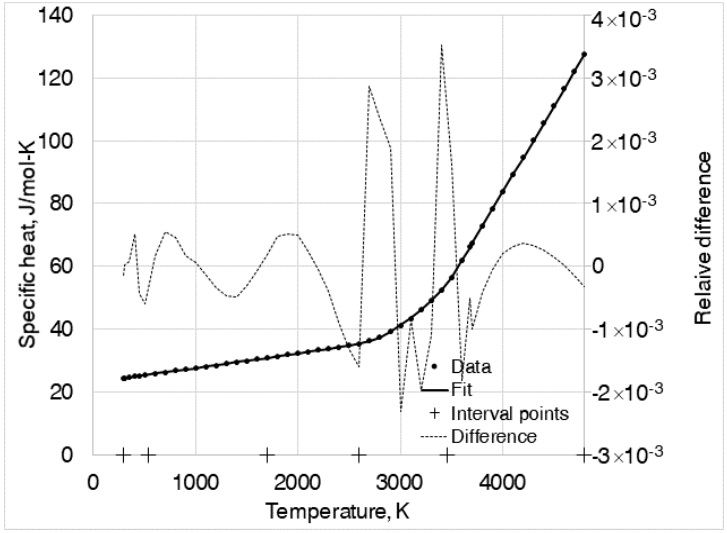


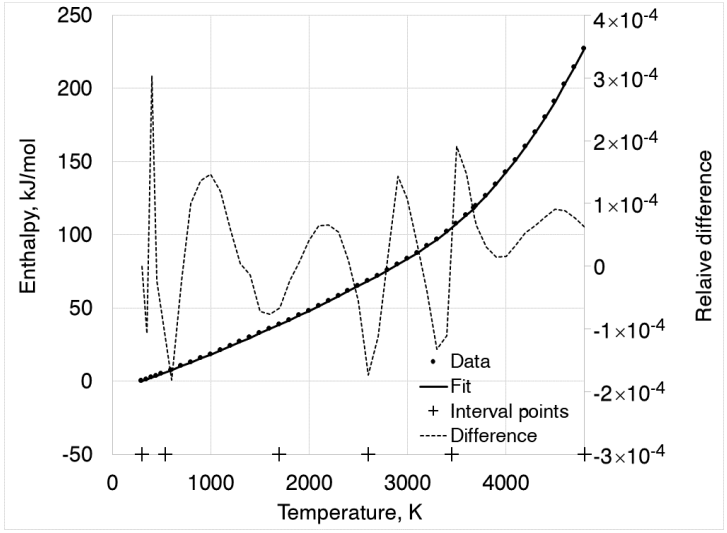


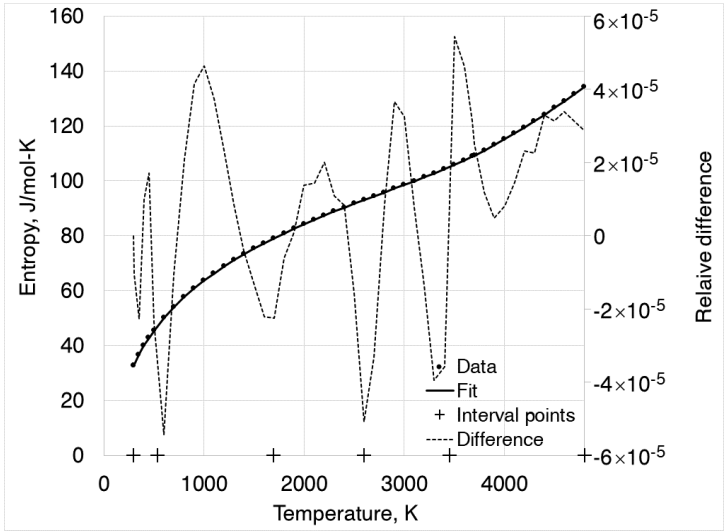


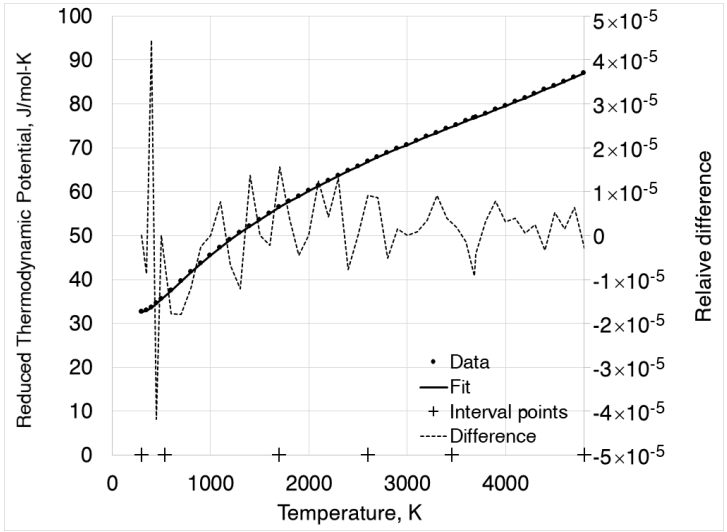
Tungsten (W) – W1(cr)

MW = 183.85 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 32.66 J/mol-K – p^o = 0.1 MPa



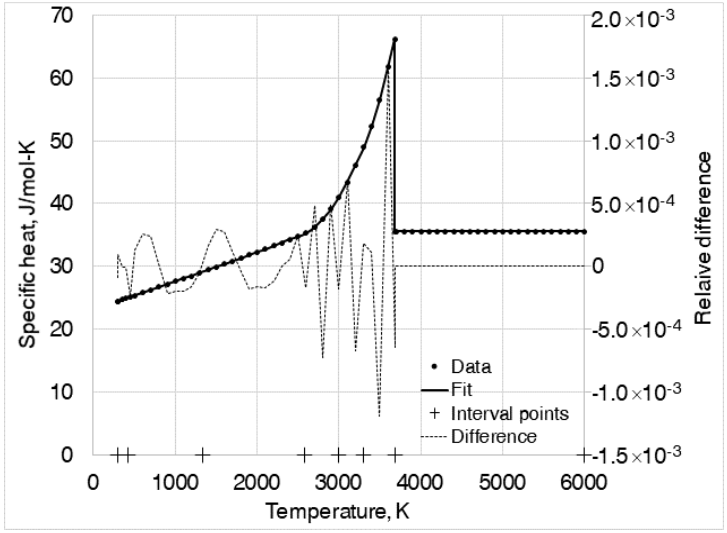


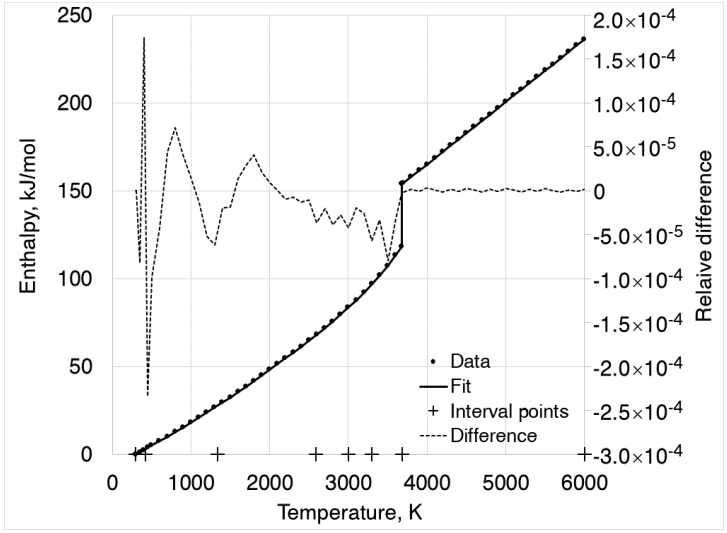


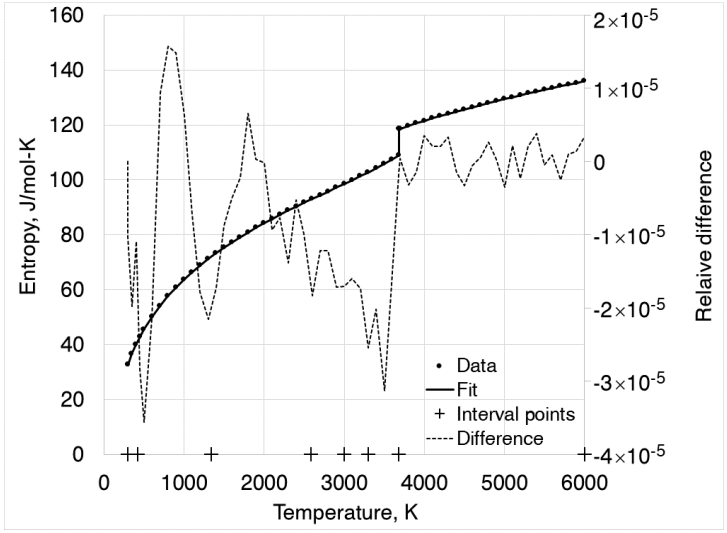


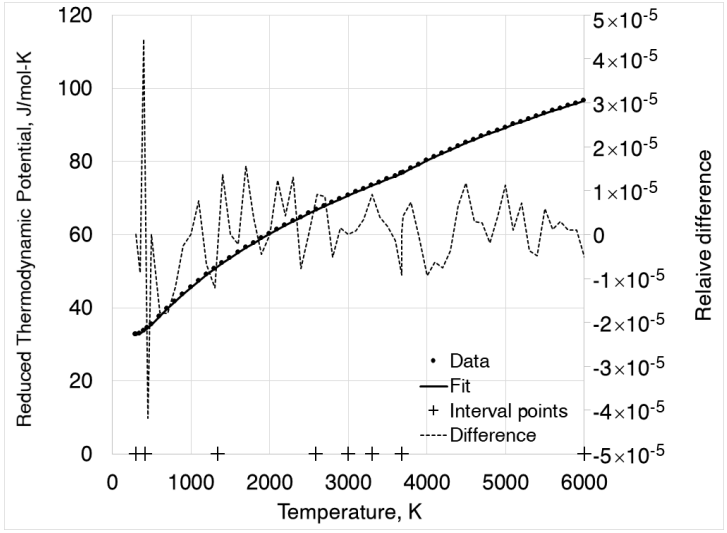
Tungsten (W) – W1(cr,l)

MW = 183.85 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 32.66 J/mol-K – p^o = 0.1 MPa



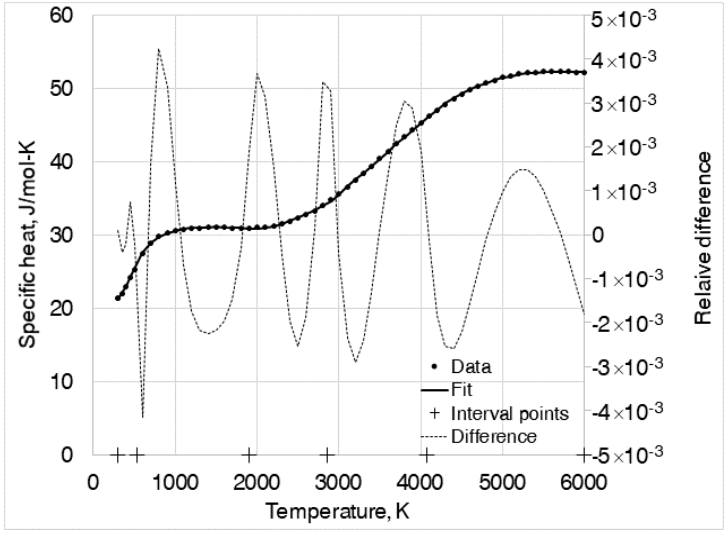


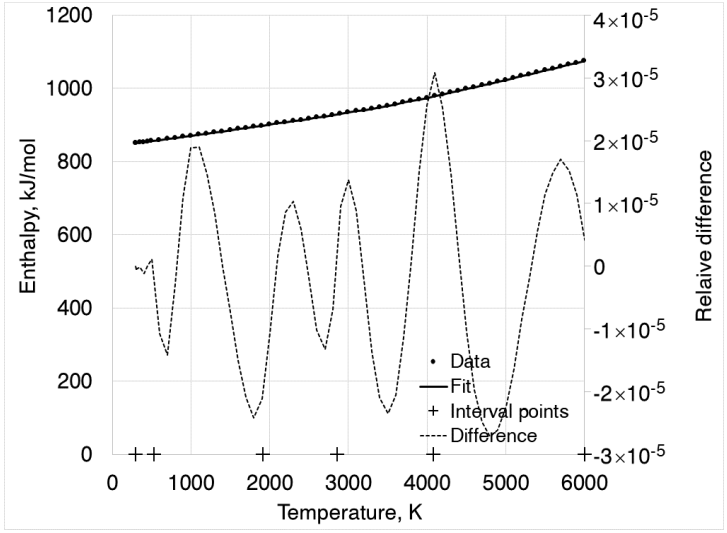


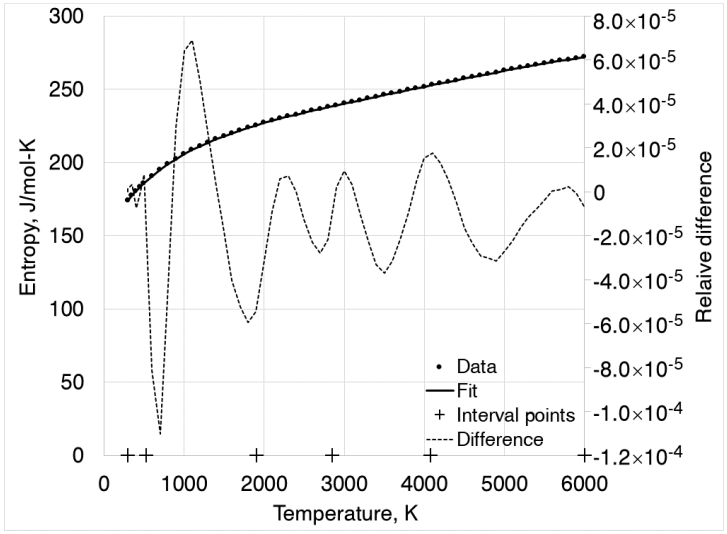


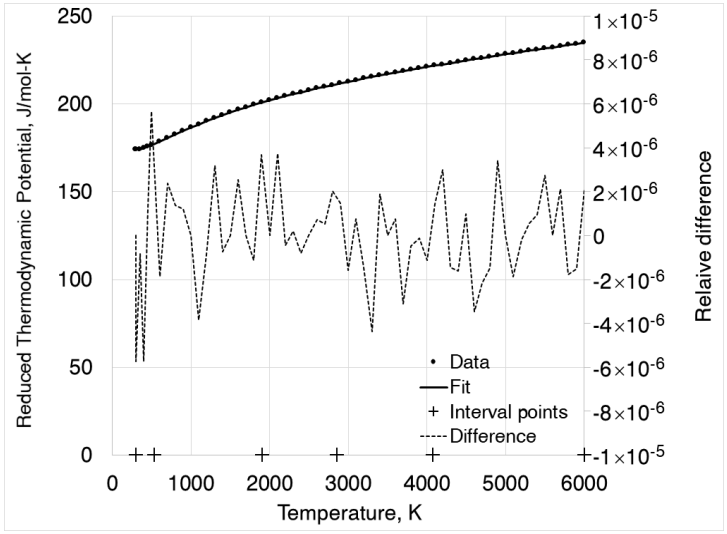
Tungsten (W) – W1(g)

MW = 183.85 g/mol – $\Delta_f H^o$(298.15 K) = 851.026 kJ/mol – S^o(298.15 K) = 173.954 J/mol-K – p^o = 0.1 MPa



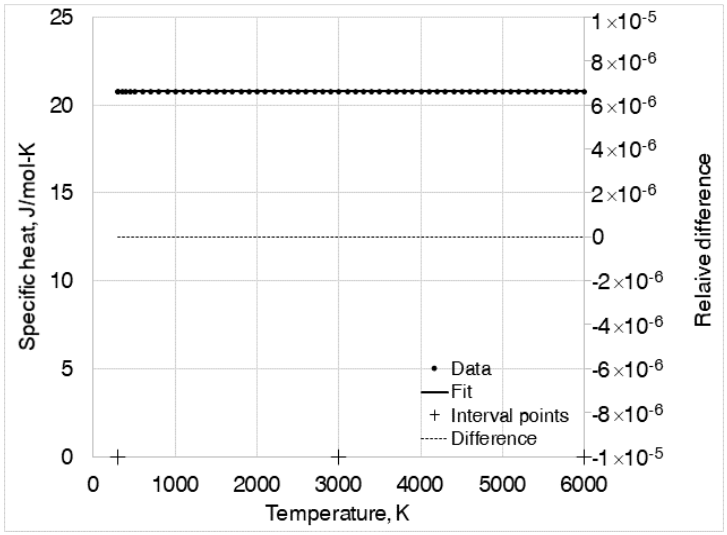


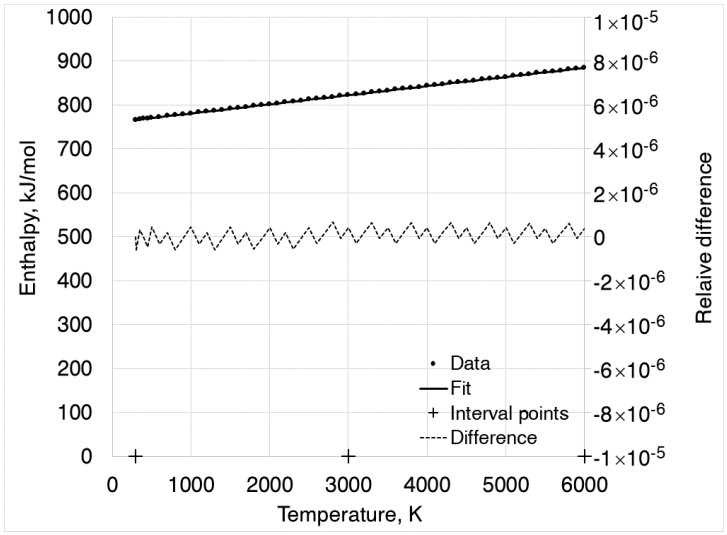


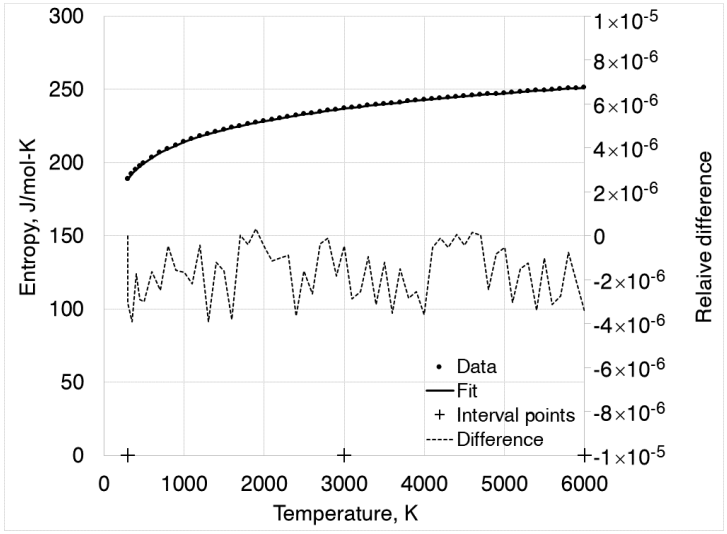


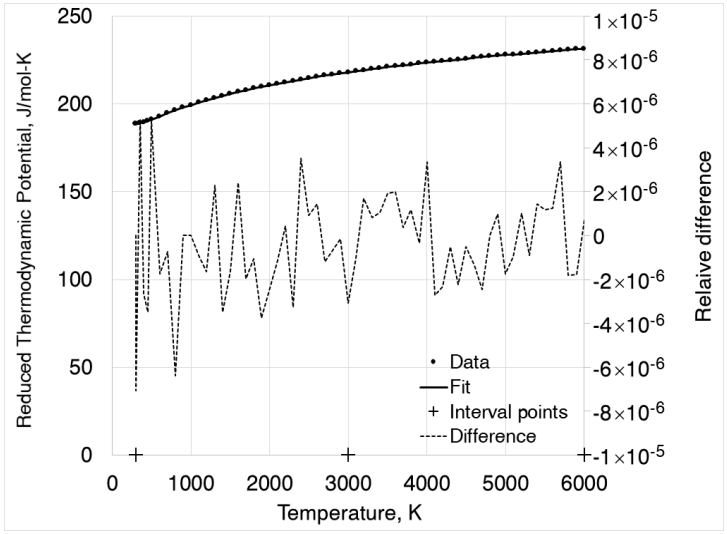
Tungsten, Ion (W^-) – W1-(g)

MW = 183.85 g/mol – $\Delta_f H^o$(298.15 K) = 766.174 kJ/mol – S^o(298.15 K) = 188.781 J/mol-K – p^o = 0.1 MPa



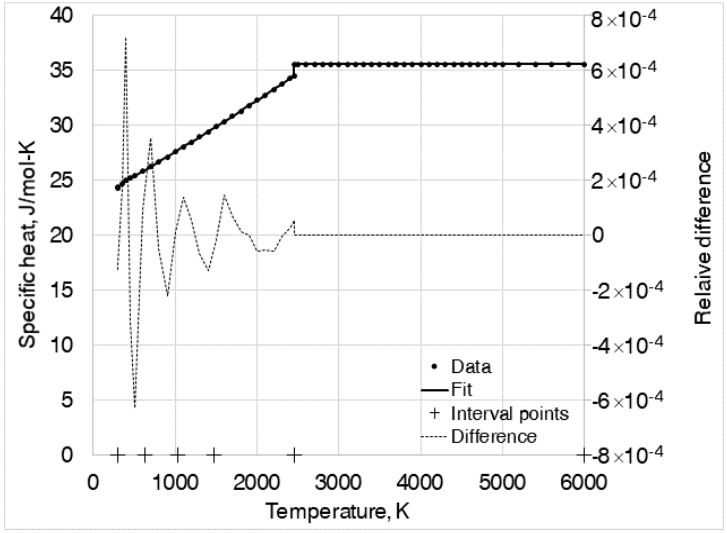


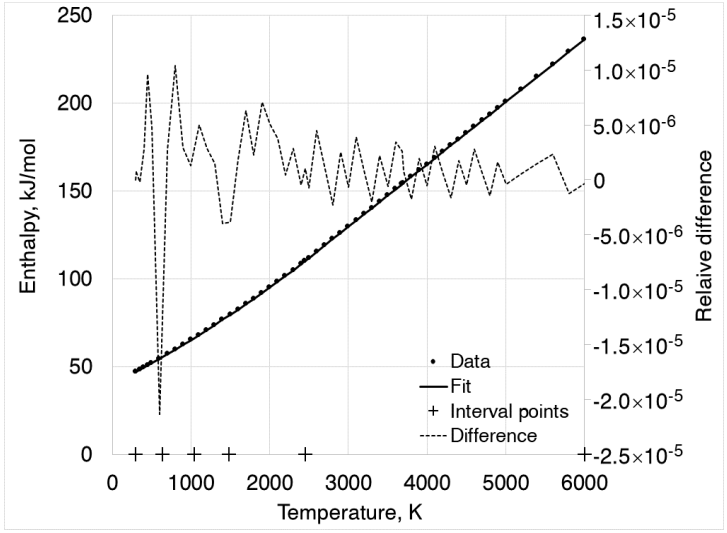


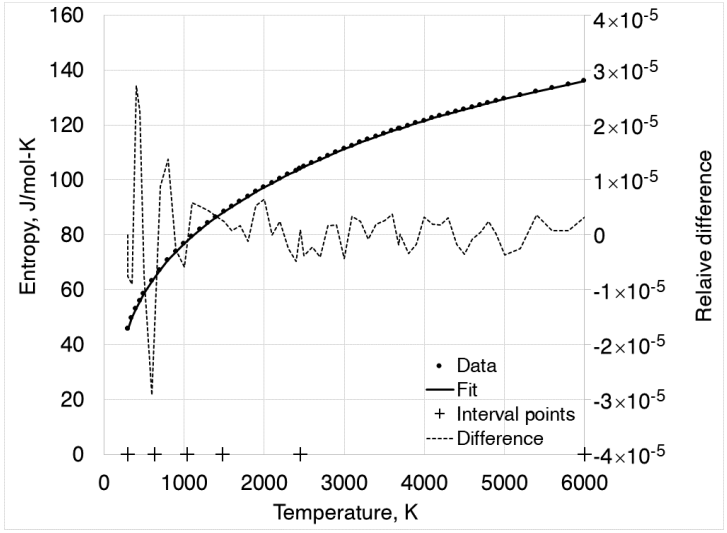


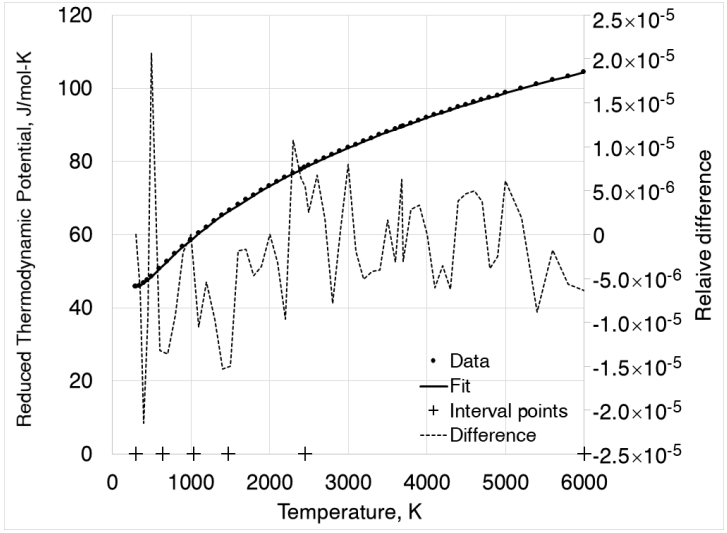
Tungsten (W) – W1(l)

MW = 183.85 g/mol – $\Delta_f H^o$(298.15 K) = 46.896 kJ/mol – S^o(298.15 K) = 45.678 J/mol-K – p^o = 0.1 MPa



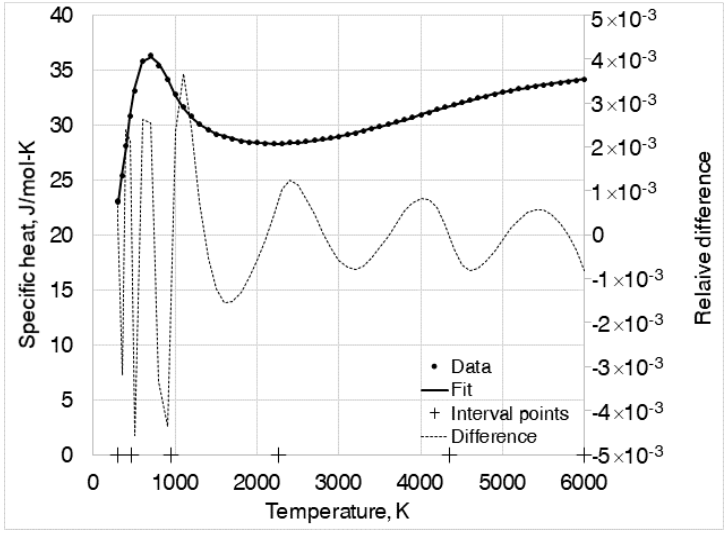


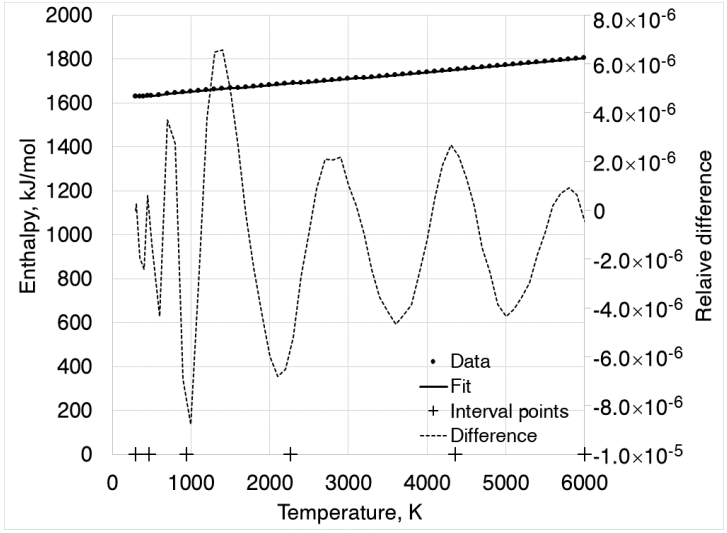


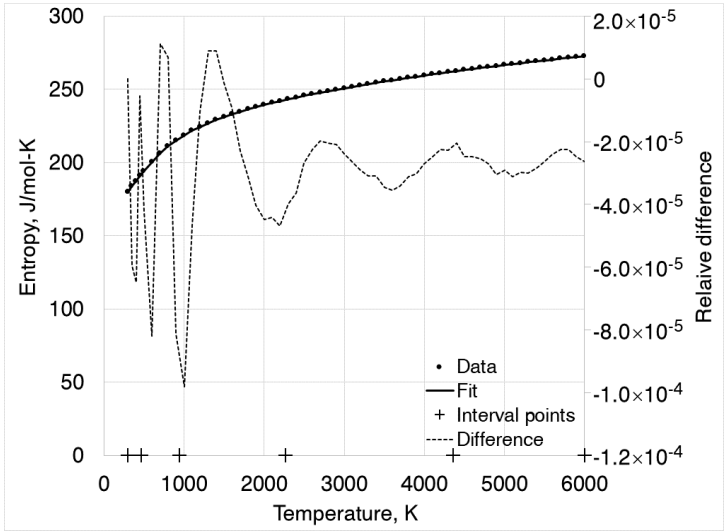


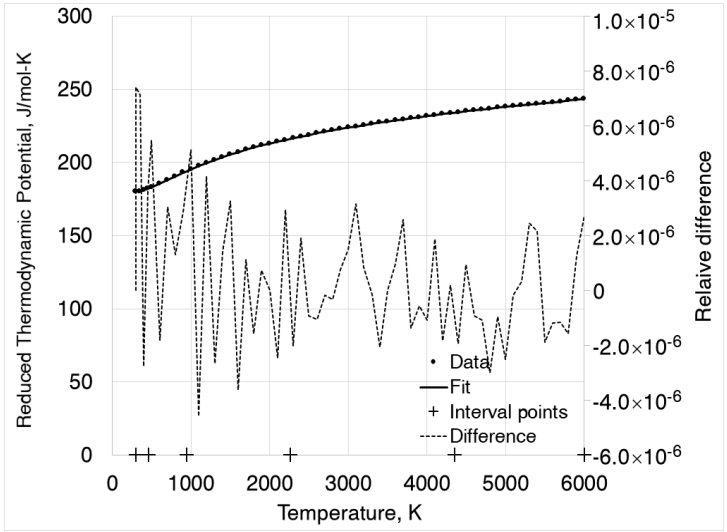
Tungsten, Ion (W^+) – W1+(g)

MW = 183.85 g/mol – $\Delta_f H^o$(298.15 K) = 1627.691 kJ/mol – S^o(298.15 K) = 179.997 J/mol-K – p^o = 0.1 MPa



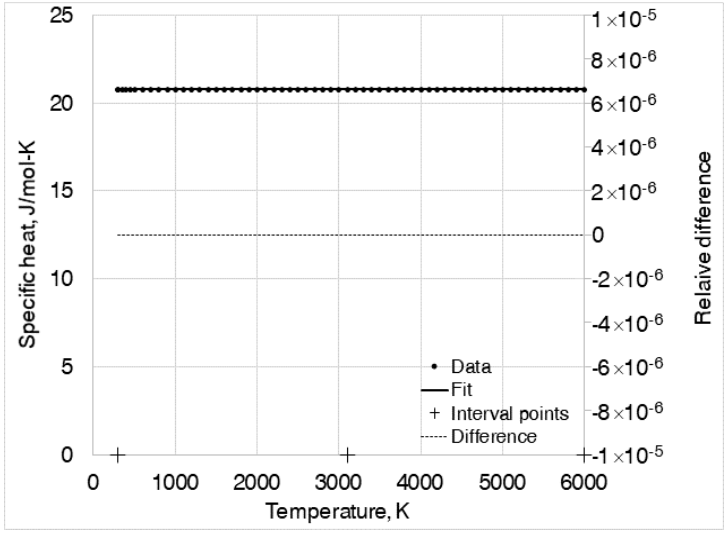


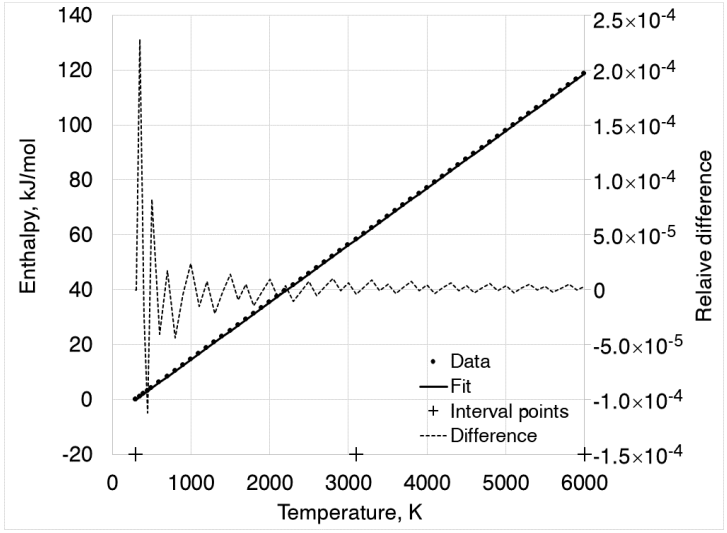


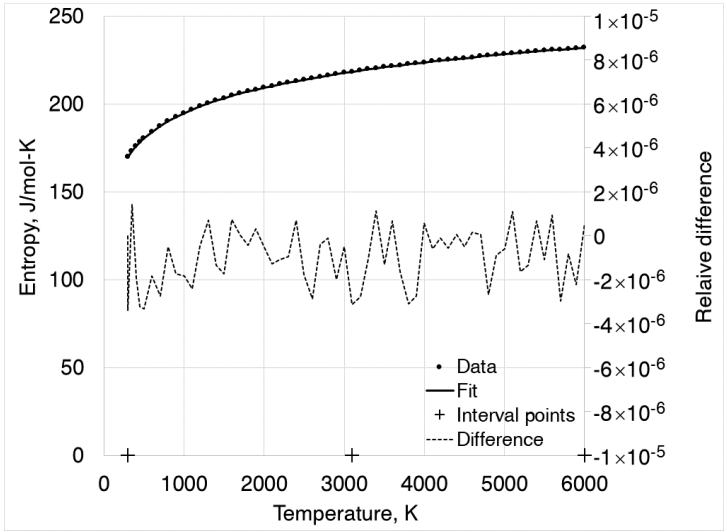


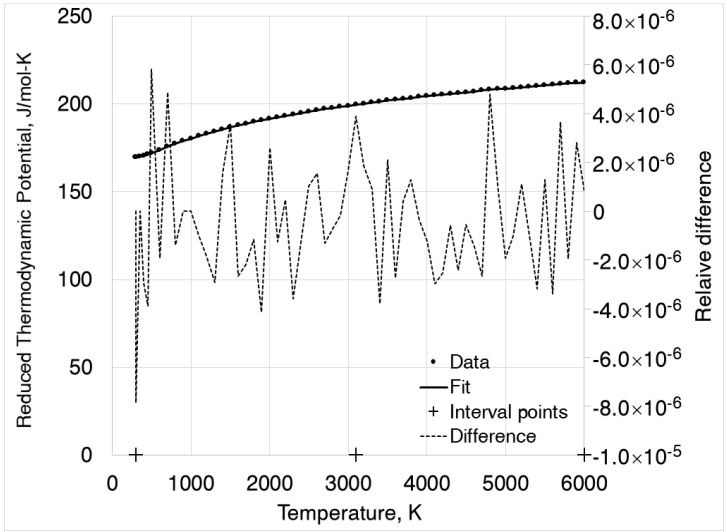
Xenon (Xe) – Xe1(ref)

MW = 131.29 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 169.684 J/mol-K – p^o = 0.1 MPa



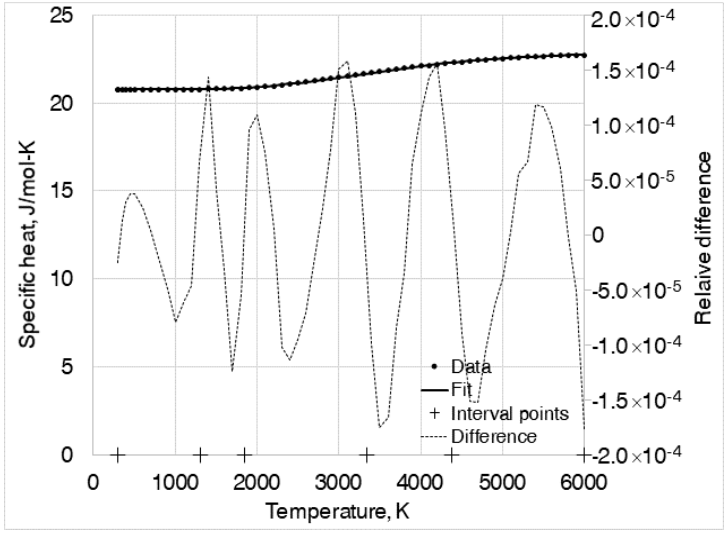


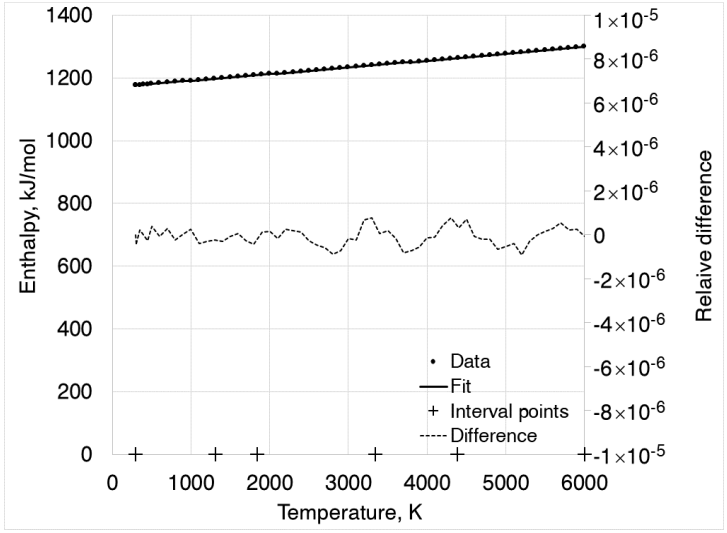


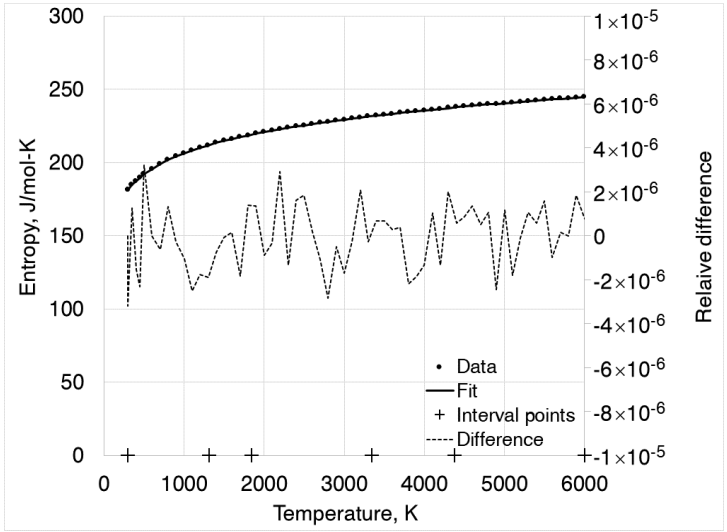


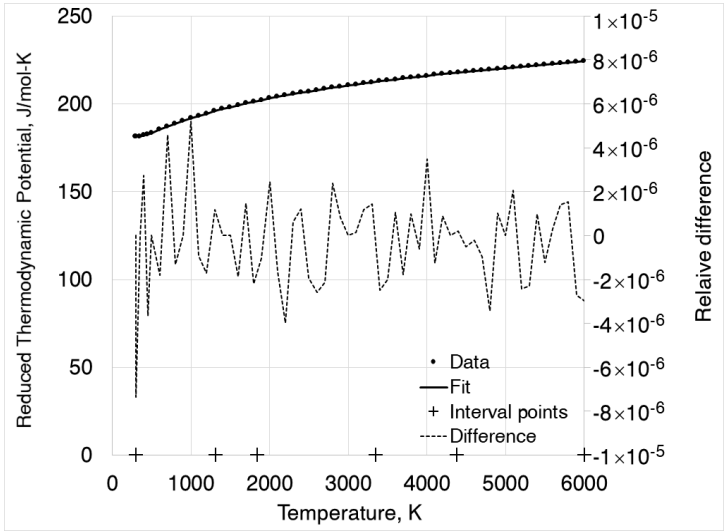
Xenon, Ion (Xe^+) – Xe1+(g)

MW = 131.29 g/mol – $\Delta_f H^o$(298.15 K) = 1176.552 kJ/mol – S^o(298.15 K) = 181.21 J/mol-K – p^o = 0.1 MPa



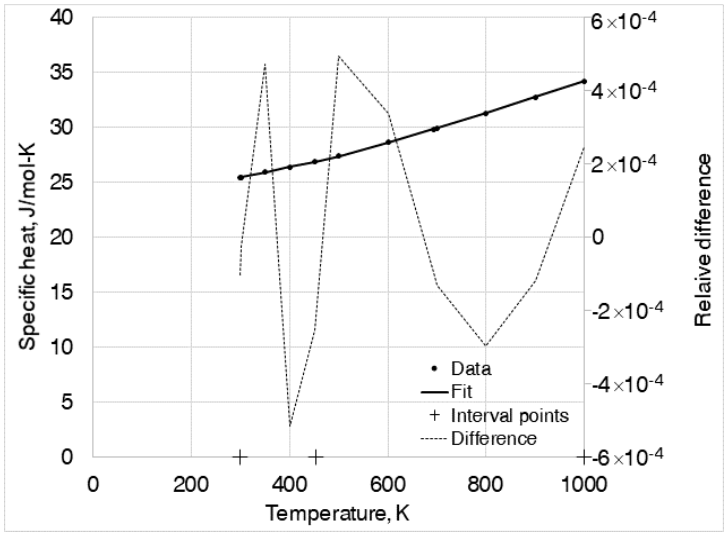


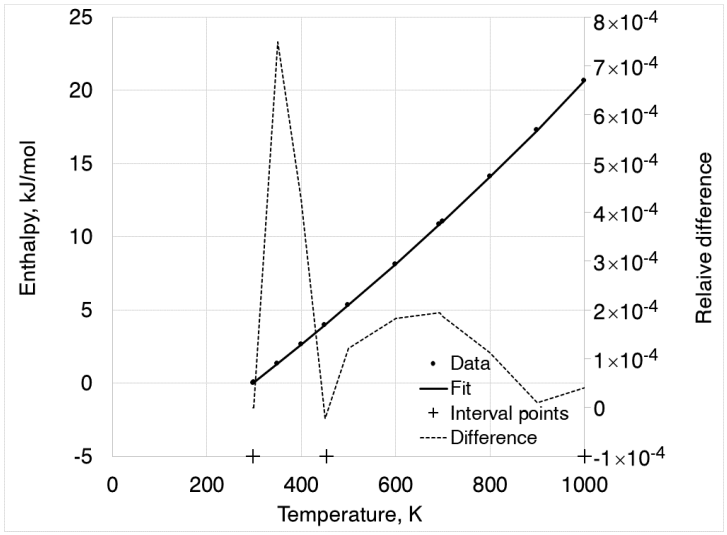


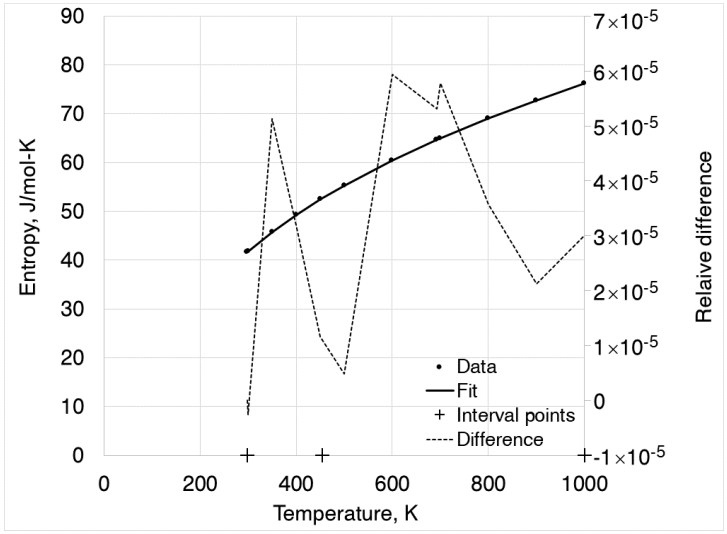


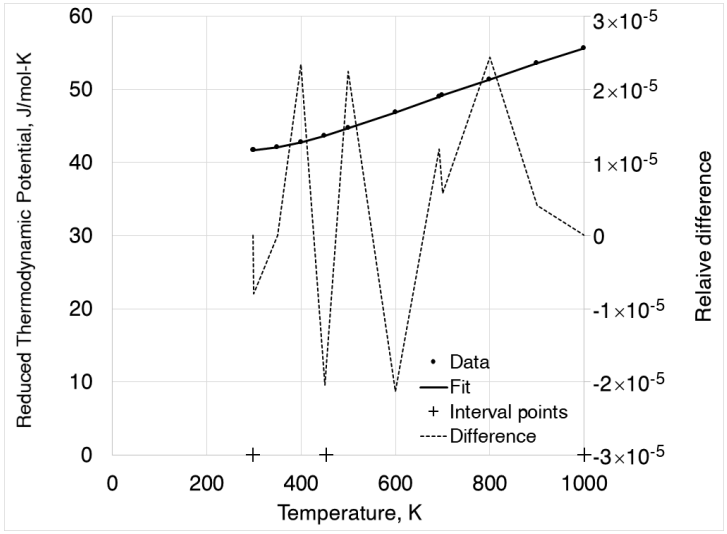
Zinc (Zn) – Zn1(cr)

MW = 65.38 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 41.717 J/mol-K – p^o = 0.1 MPa



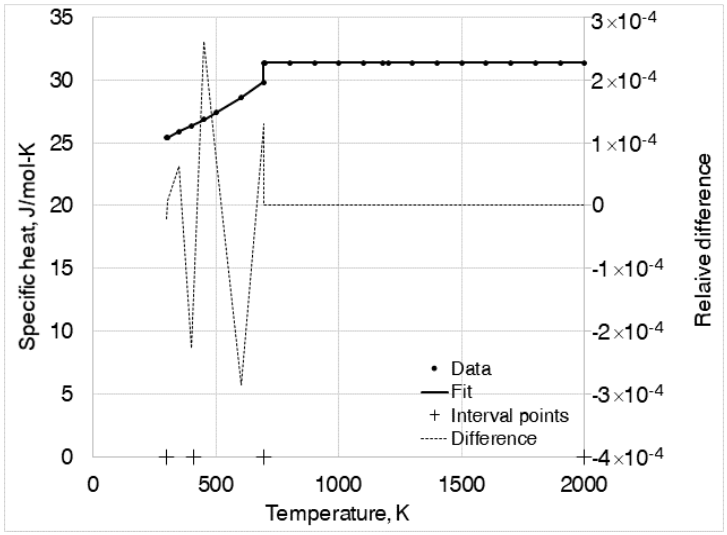


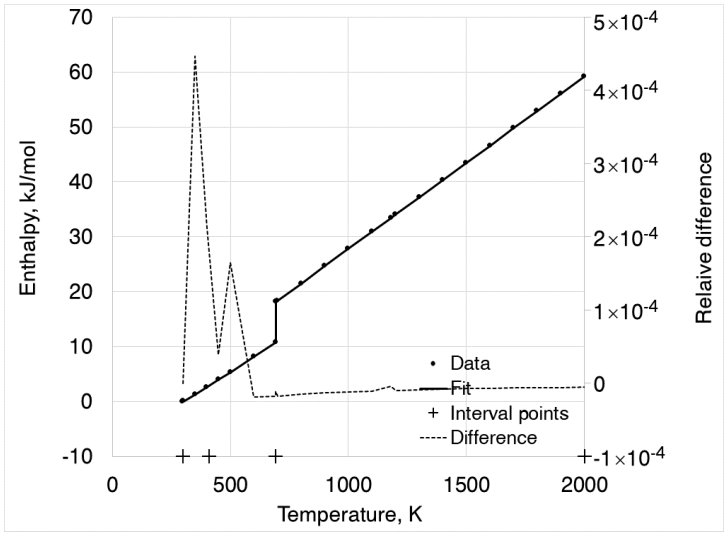


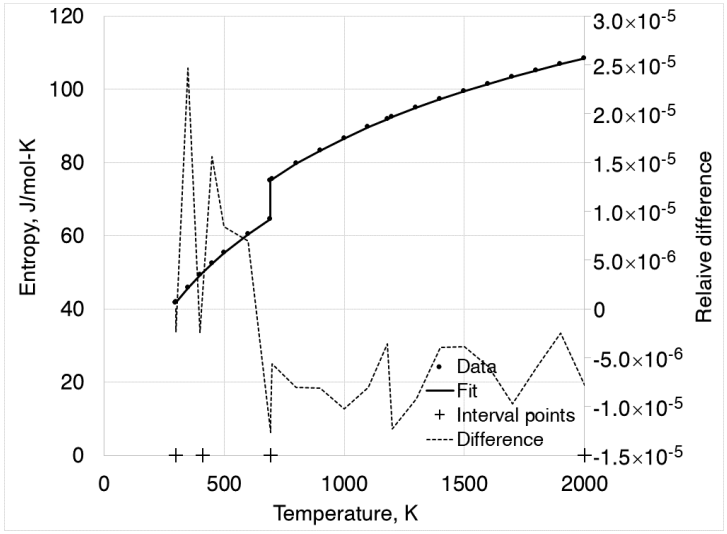


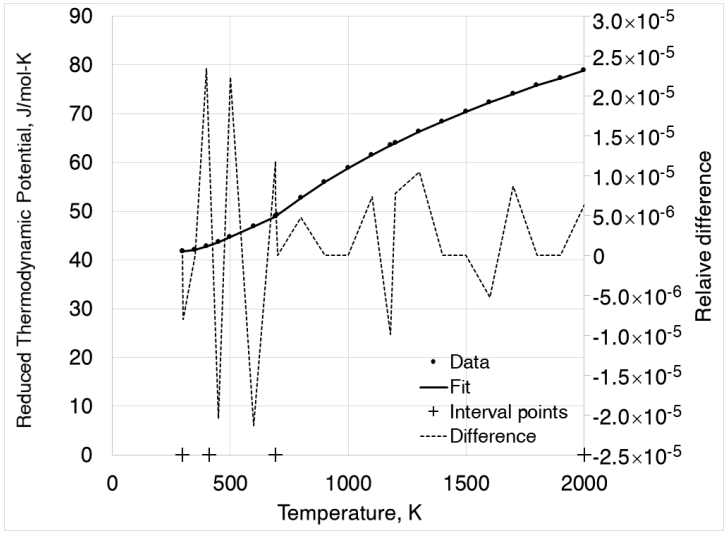
Zinc (Zn) – Zn1(cr,l)

MW = 65.38 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 41.717 J/mol-K – p^o = 0.1 MPa



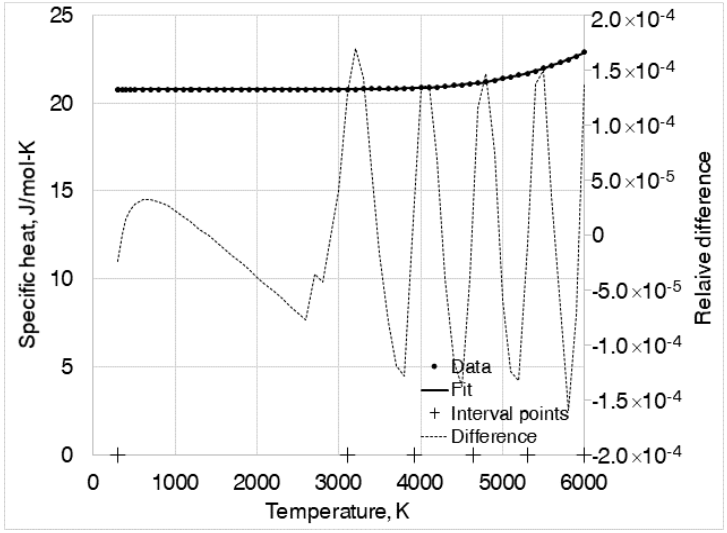


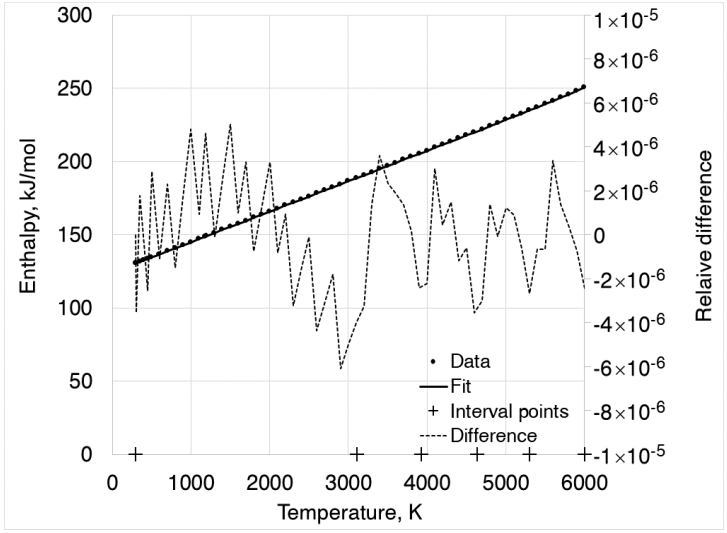


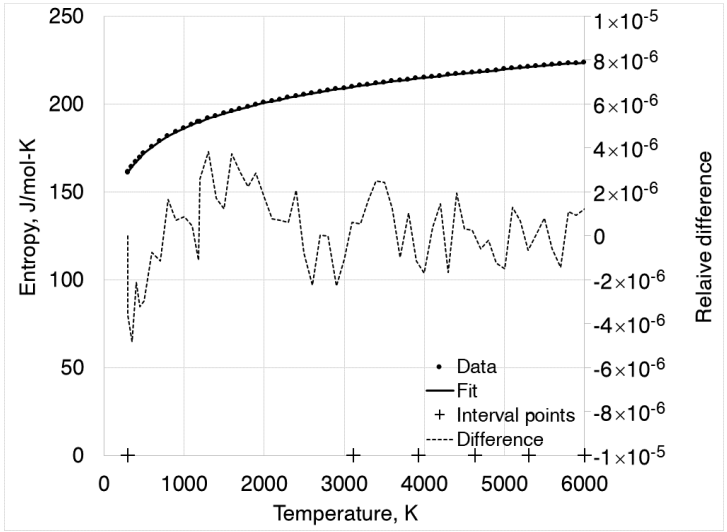


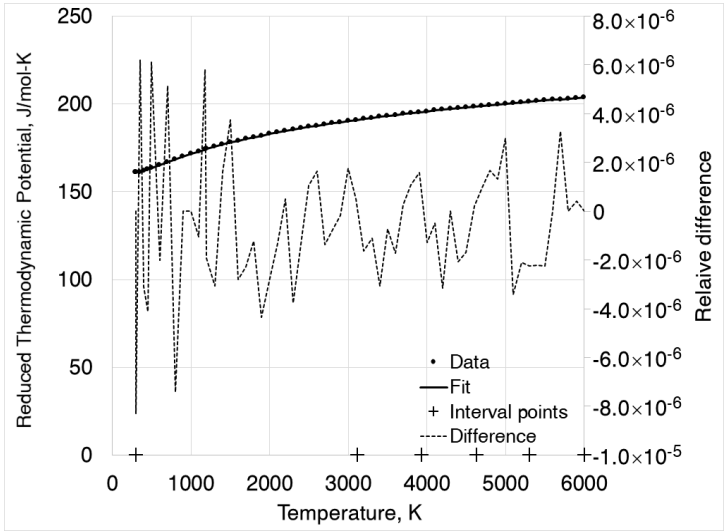
Zinc (Zn) – Zn1(g)

MW = 65.38 g/mol – $\Delta_f H^o$(298.15 K) = 130.42 kJ/mol – S^o(298.15 K) = 160.989 J/mol-K – p^o = 0.1 MPa



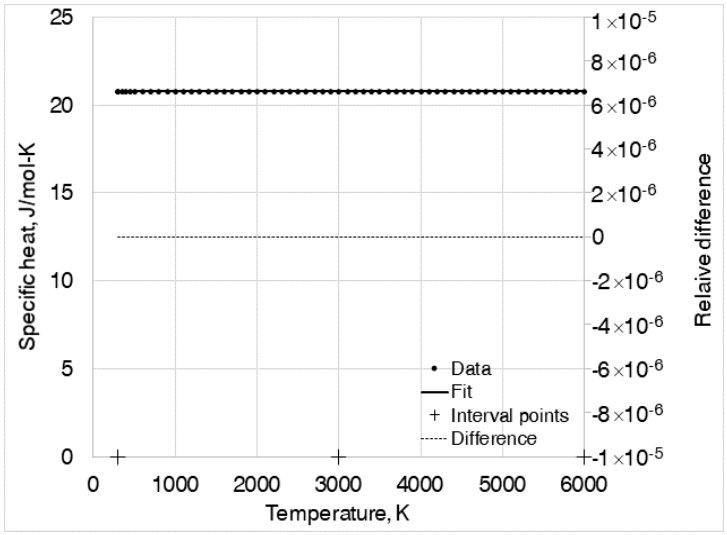


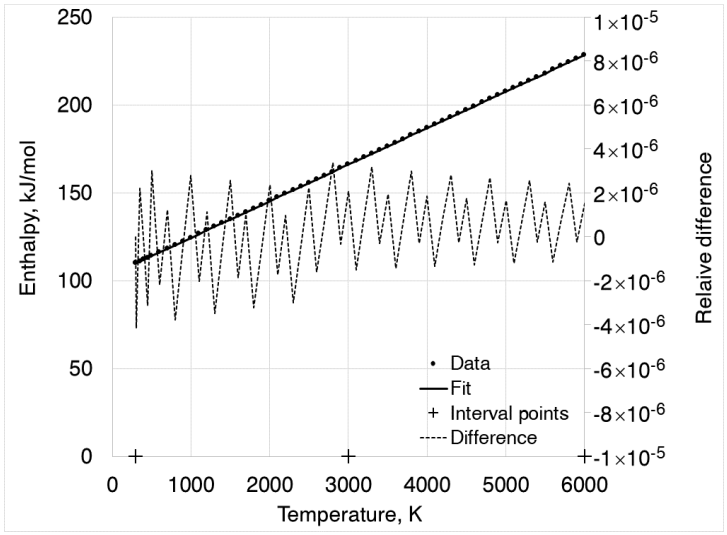


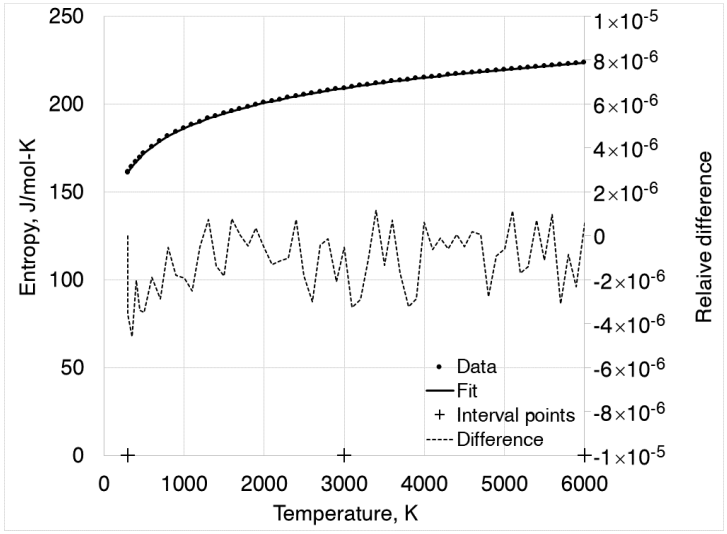


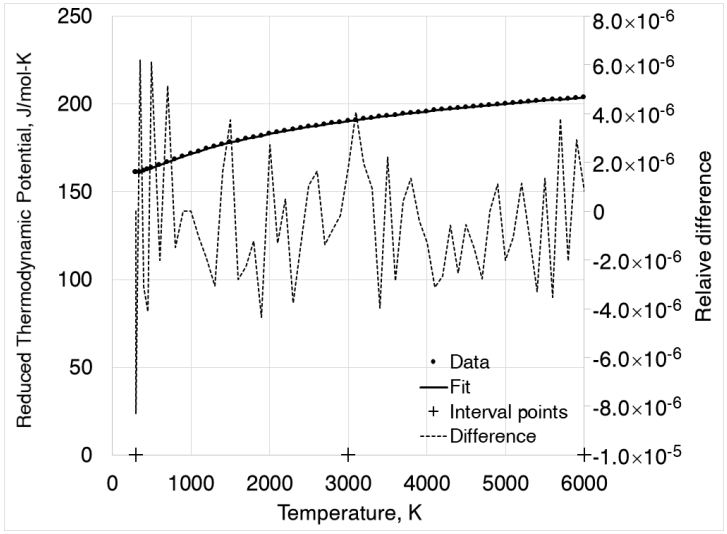
Zinc, Ion (Zn^-) – Zn1-(g)

MW = 65.38 g/mol – $\Delta_f H^o$(298.15 K) = 109.87 kJ/mol – S^o(298.15 K) = 160.989 J/mol-K – p^o = 0.1 MPa



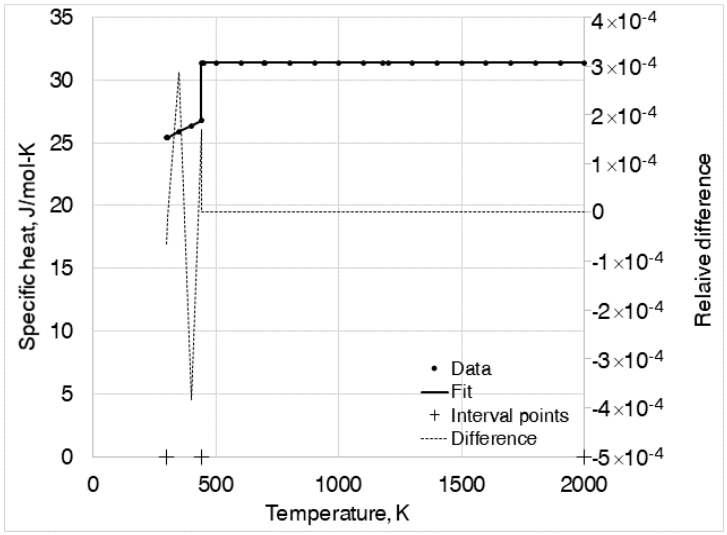


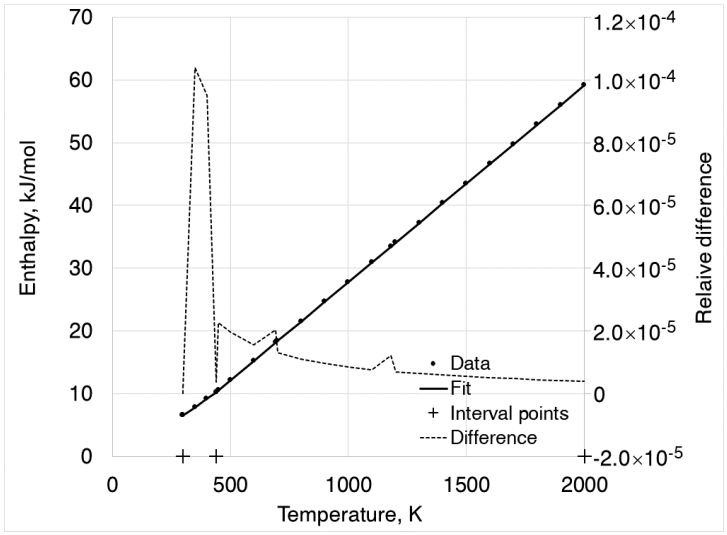


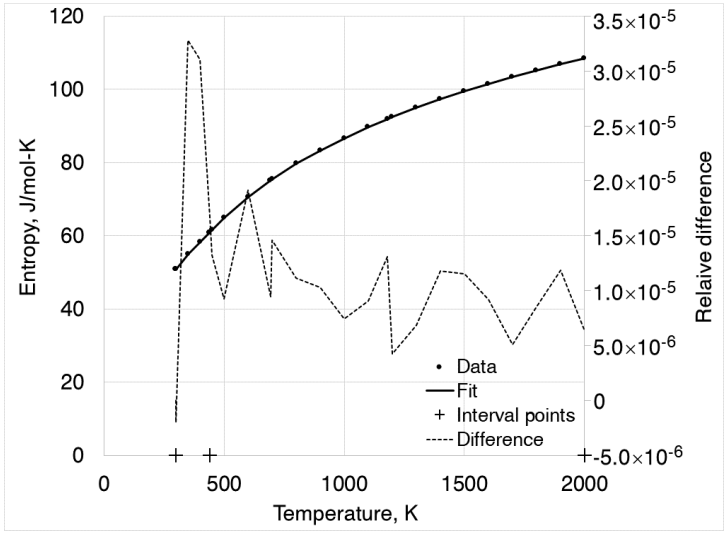


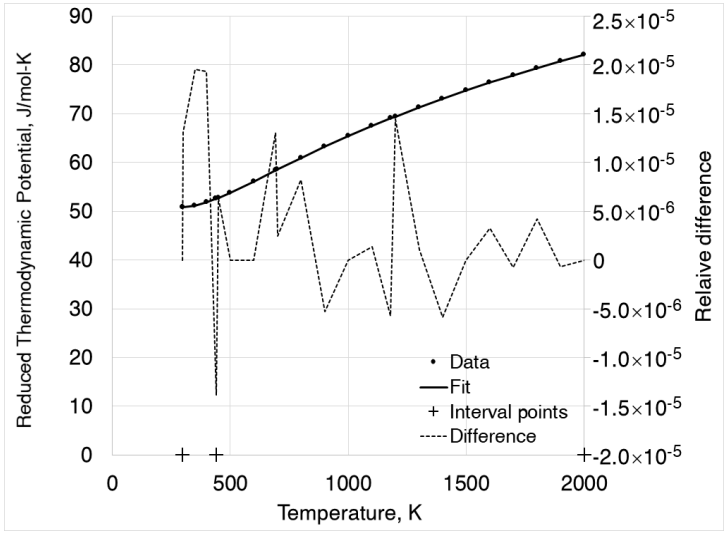
Zinc (Zn) – Zn1(l)

MW = 65.38 g/mol – $\Delta_f H^o$(298.15 K) = 6.519 kJ/mol – S^o(298.15 K) = 50.792 J/mol-K – p^o = 0.1 MPa



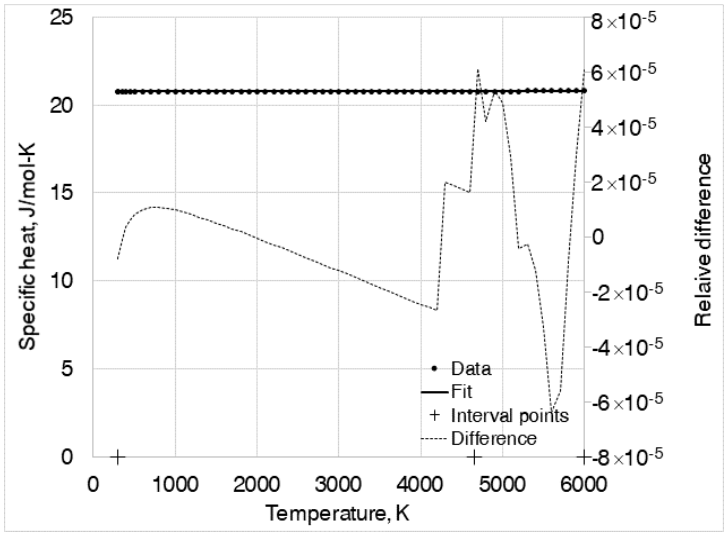


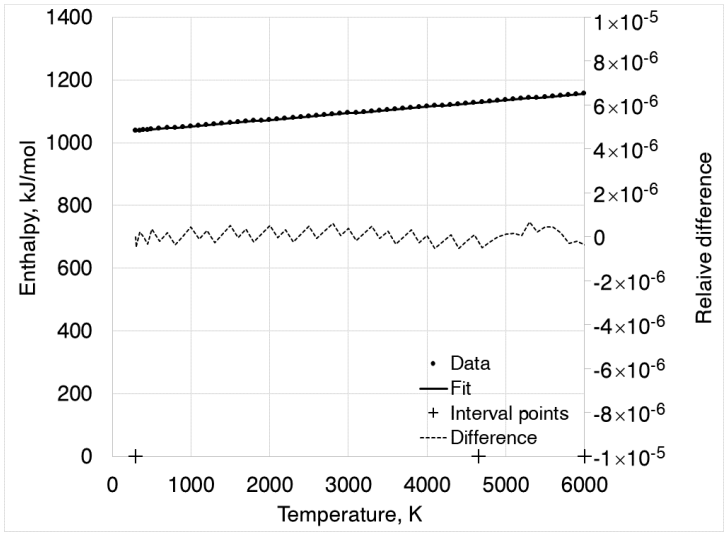


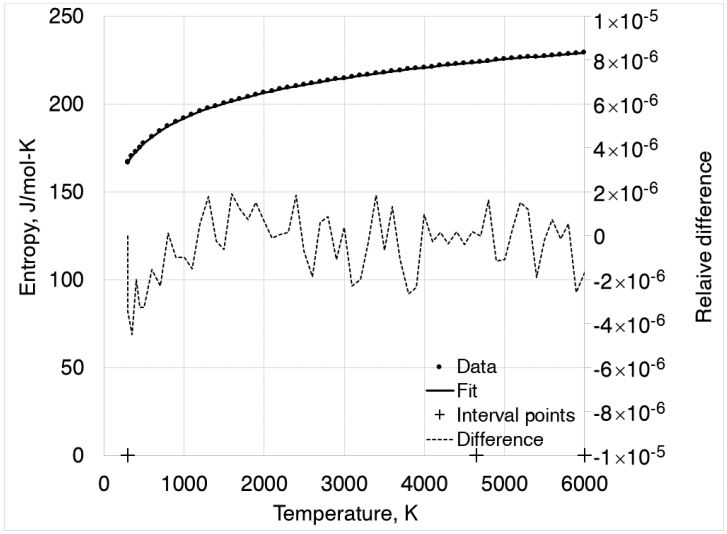


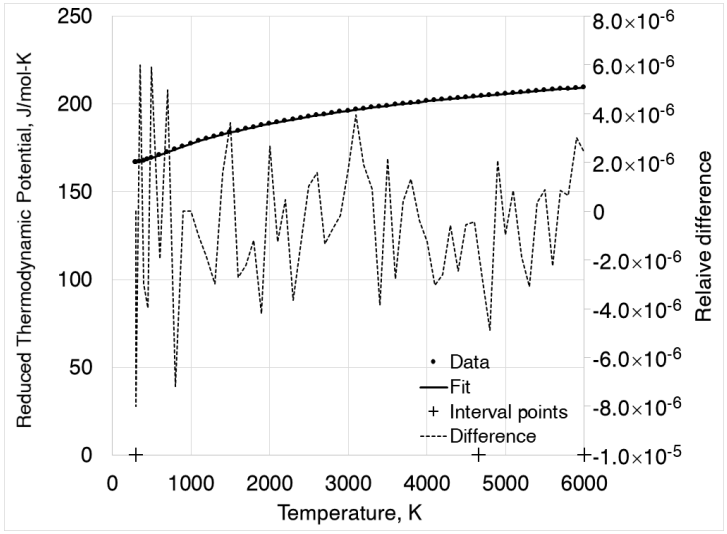
Zinc, Ion (Zn^+) – Zn1+(g)

MW = 65.38 g/mol – $\Delta_f H^o$(298.15 K) = 1037.336 kJ/mol – S^o(298.15 K) = 166.751 J/mol-K – p^o = 0.1 MPa



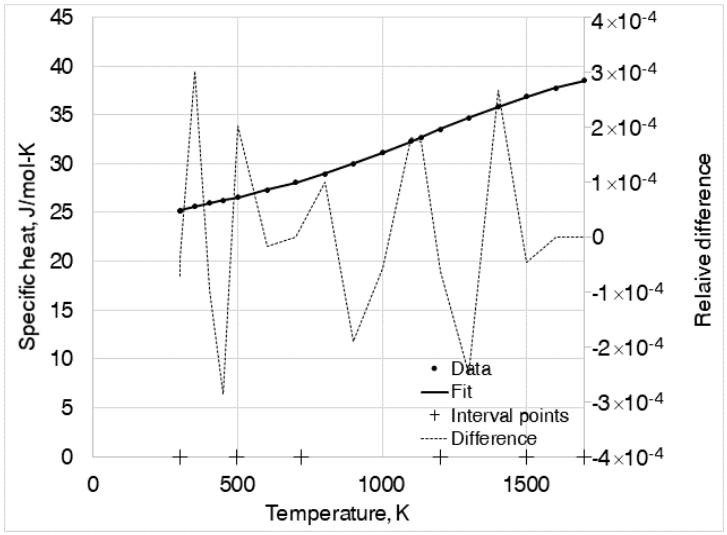


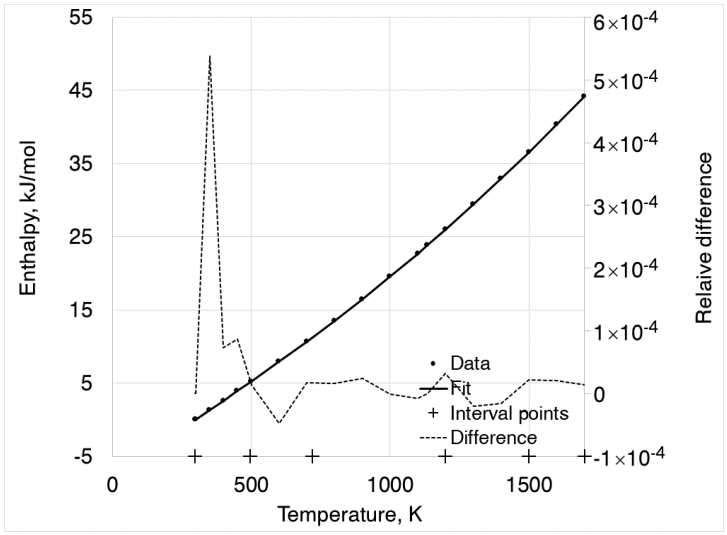


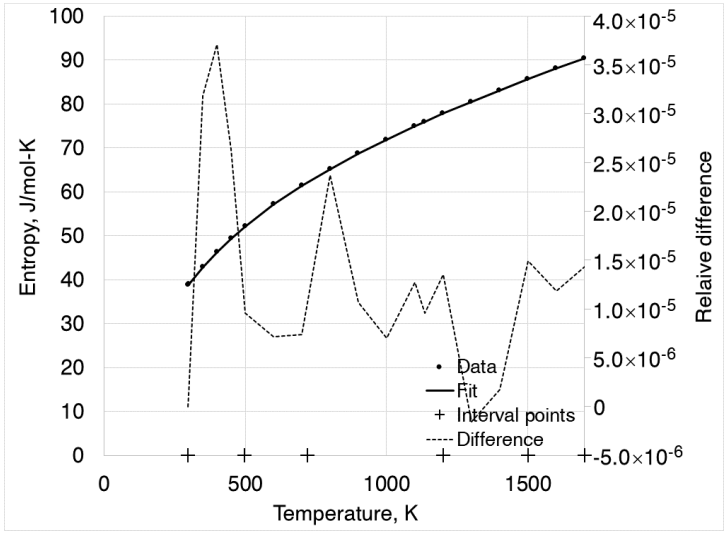


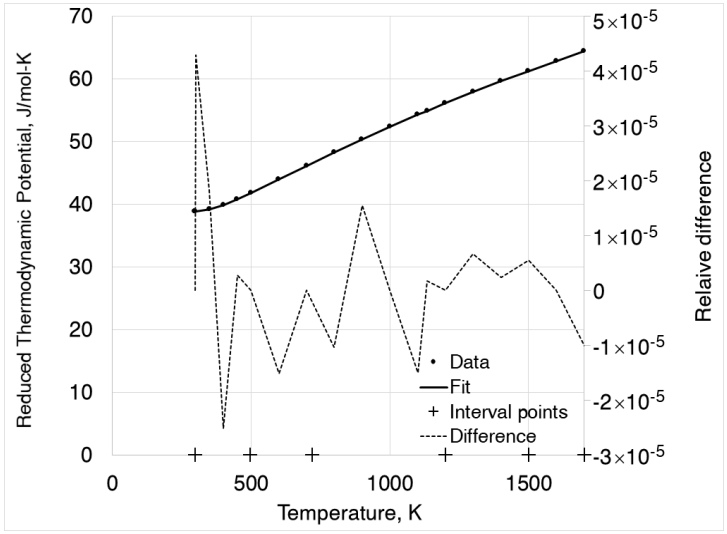
Zirconium, Alpha (Zr) – Zr1(cr)_01

MW = 91.22 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 38.869 J/mol-K – p^o = 0.1 MPa



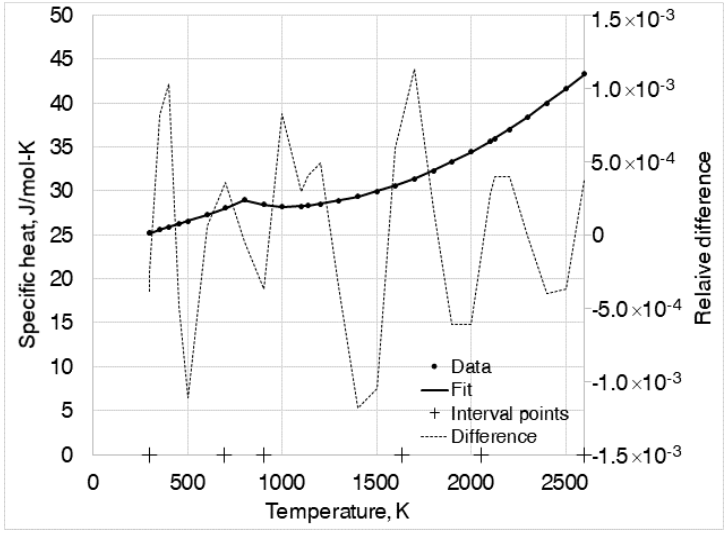


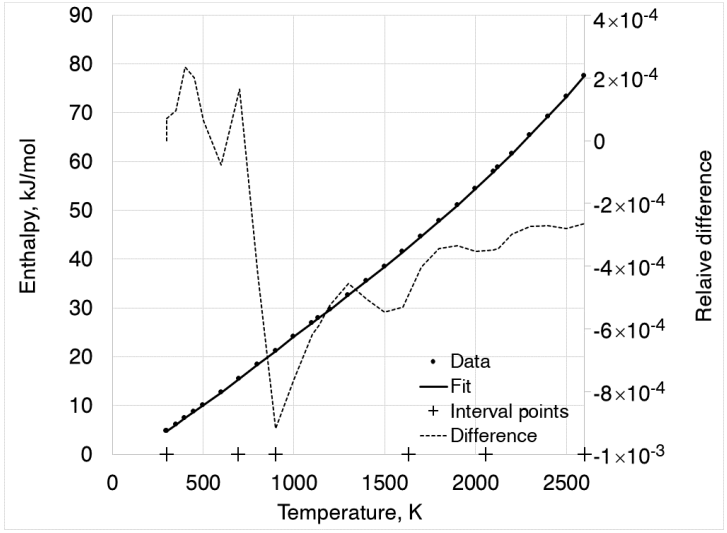


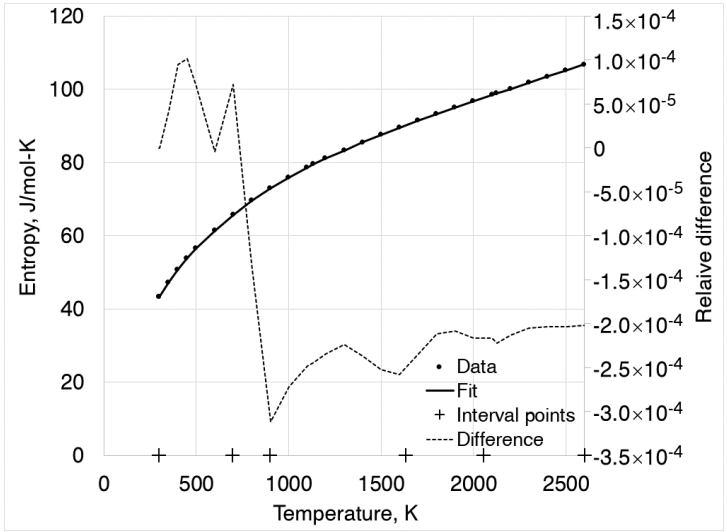


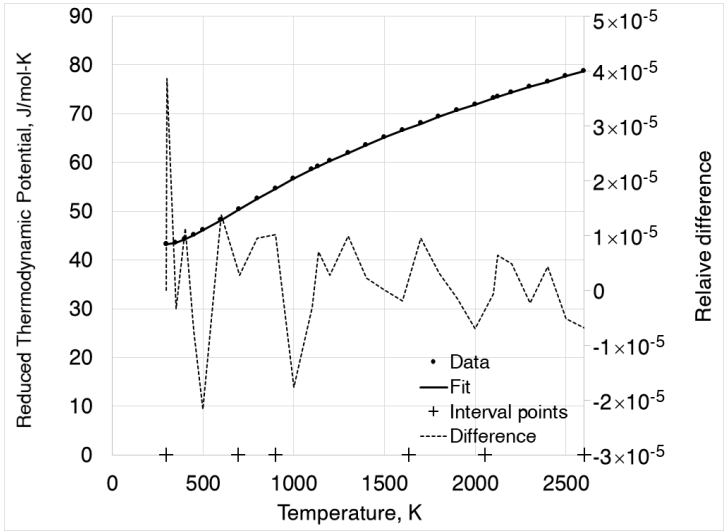
Zirconium, Beta (Zr) – Zr1(cr)_02

MW = 91.22 g/mol – $\Delta_f H^o$(298.15 K) = 4.81 kJ/mol – S^o(298.15 K) = 43.192 J/mol-K – p^o = 0.1 MPa



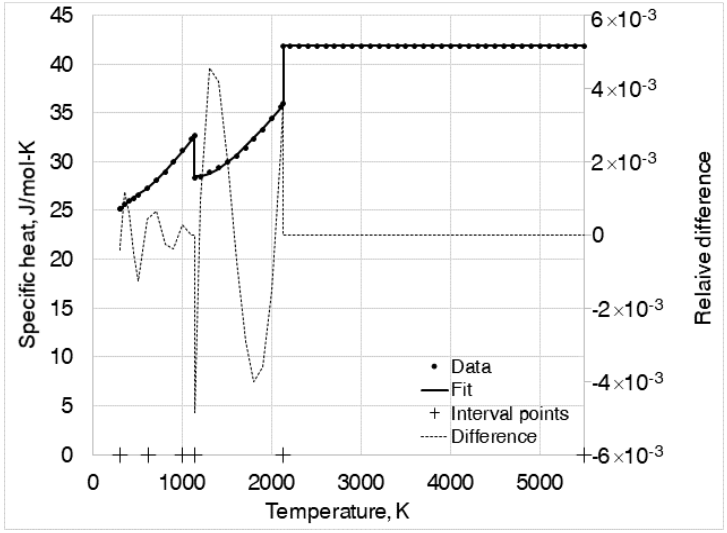


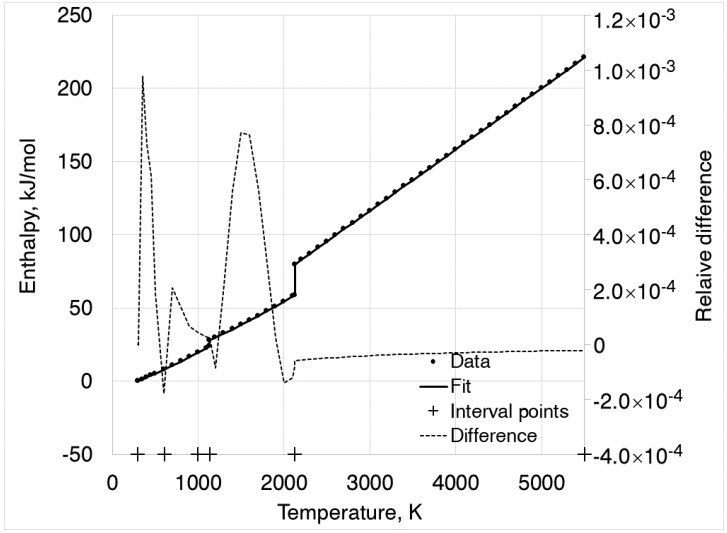


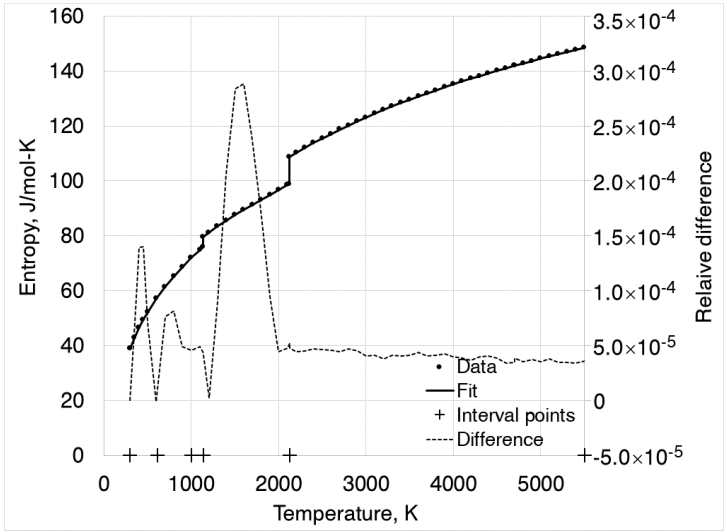


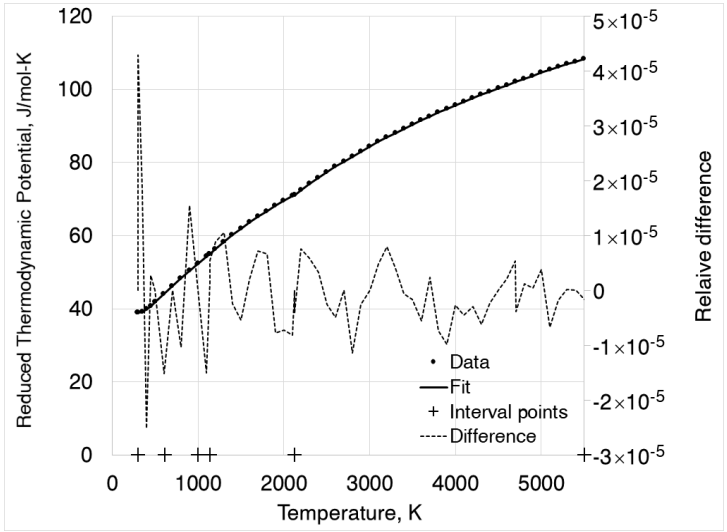
Zirconium (Zr) – Zr1(cr,l)

MW = 91.22 g/mol – $\Delta_f H^o$(298.15 K) = 0 kJ/mol – S^o(298.15 K) = 38.869 J/mol-K – p^o = 0.1 MPa



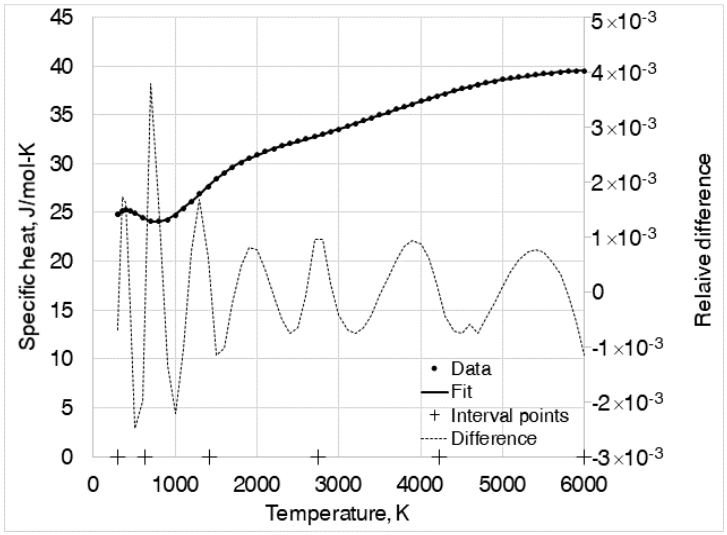


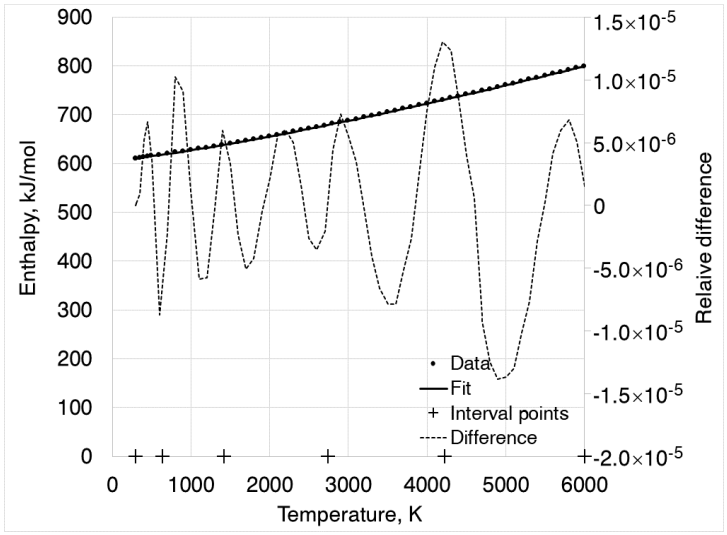


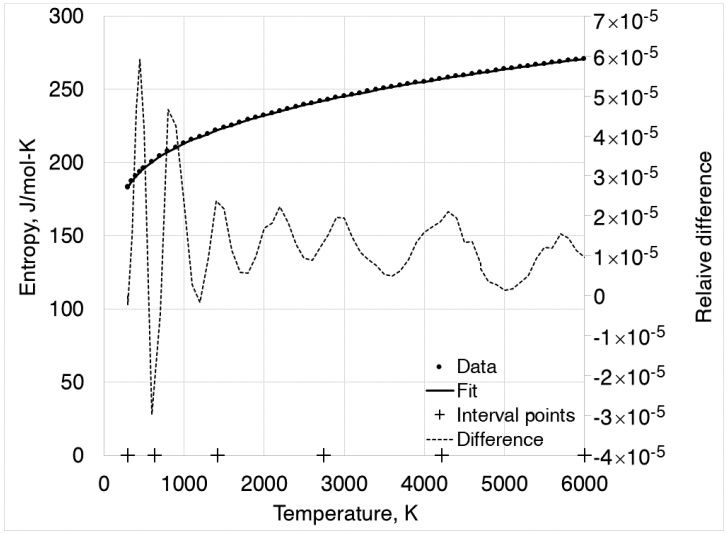


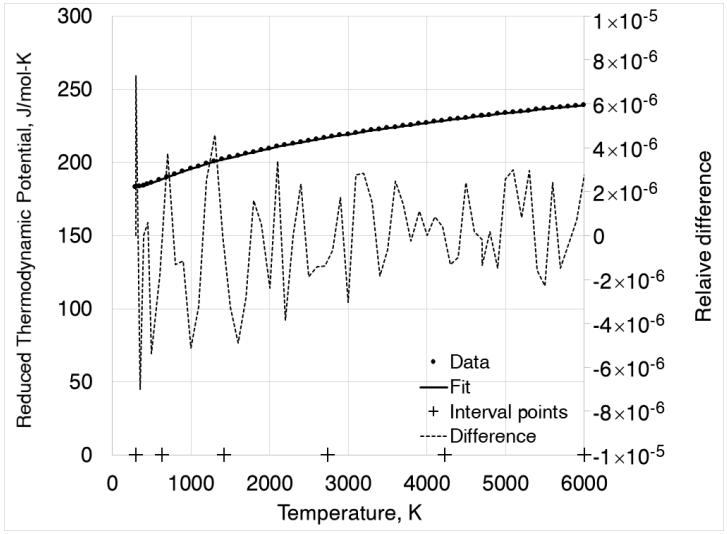
Zirconium (Zr) – Zr1(g)

MW = 91.22 g/mol – $\Delta_f H^o$(298.15 K) = 610.027 kJ/mol – S^o(298.15 K) = 183.027 J/mol-K – p^o = 0.1 MPa



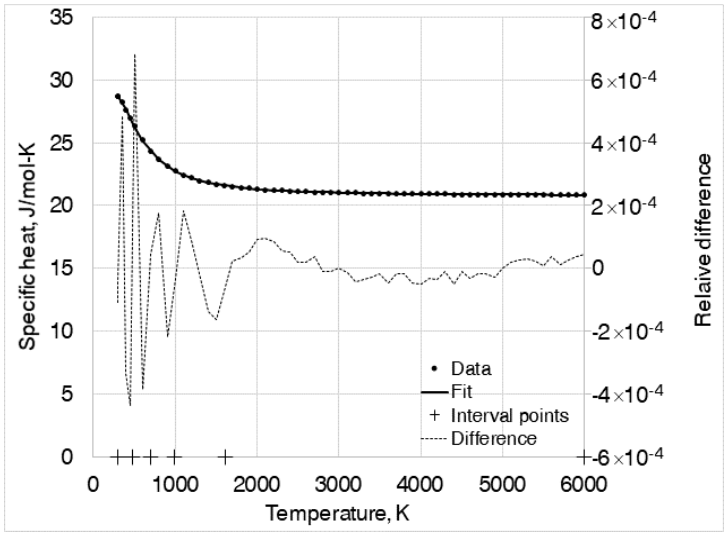


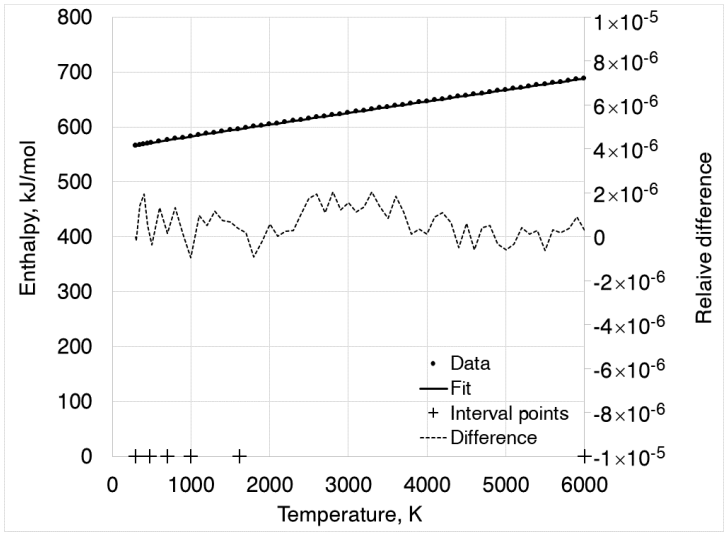


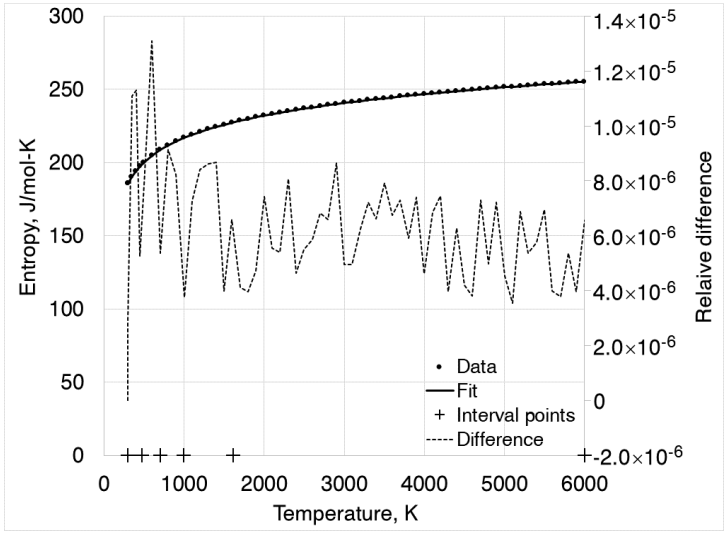


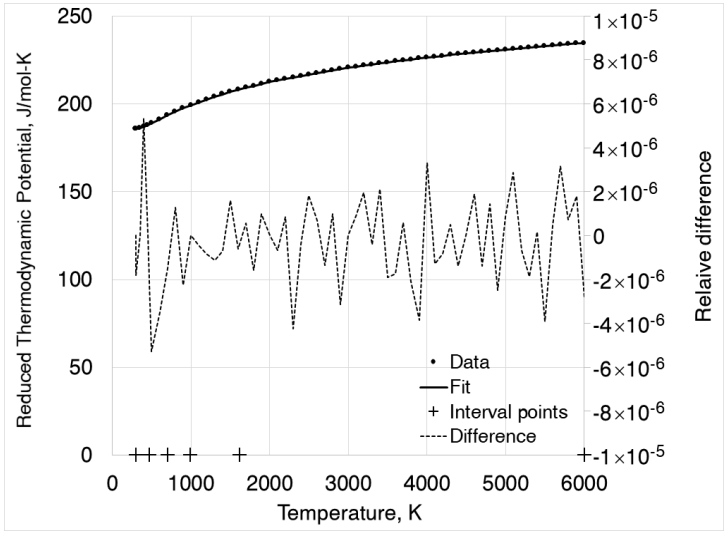
Zirconium, Ion (Zr^-) – Zr1-(g)

MW = 91.22 g/mol – $\Delta_f H^o$(298.15 K) = 564.979 kJ/mol – S^o(298.15 K) = 185.762 J/mol-K – p^o = 0.1 MPa



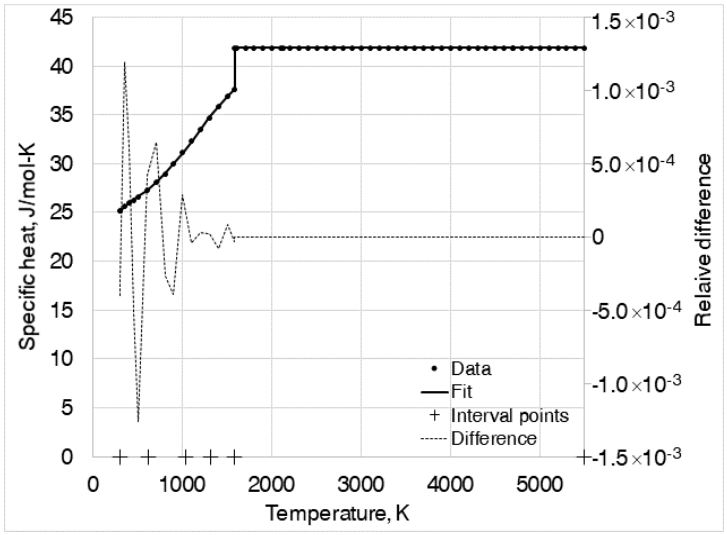


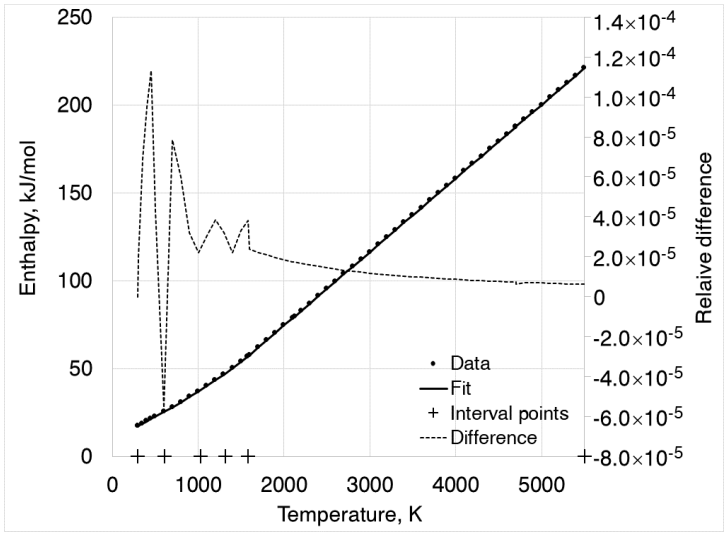


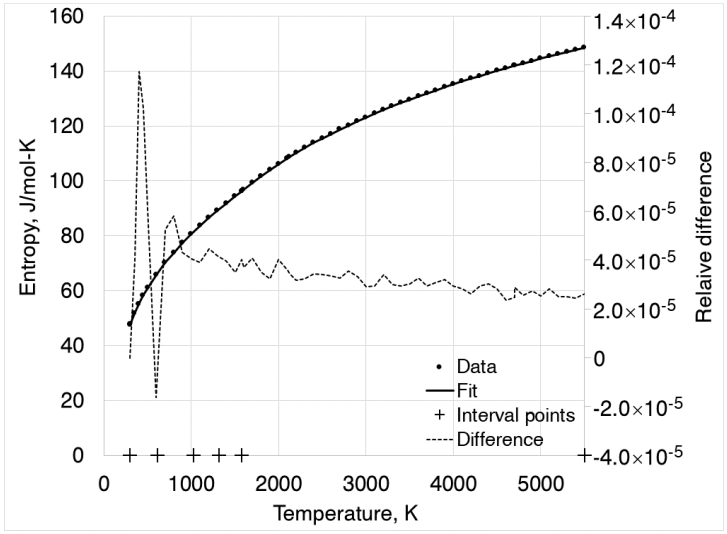


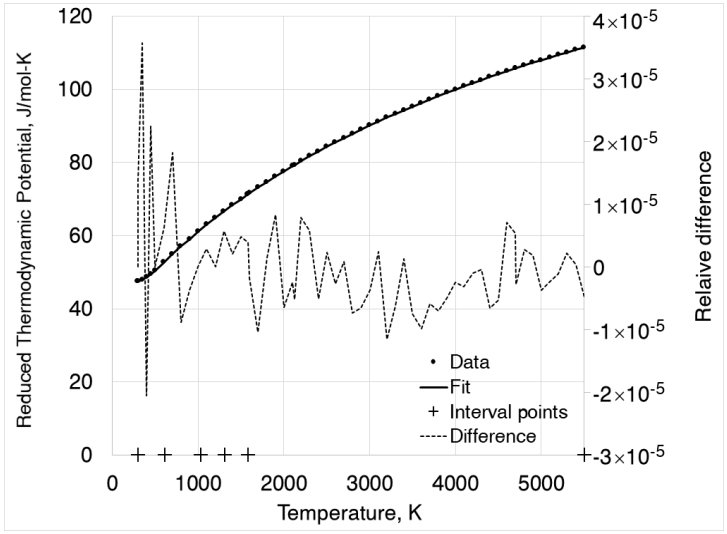
Zirconium (Zr) – Zr1(l)

MW = 91.22 g/mol – $\Delta_f H^o$(298.15 K) = 17.404 kJ/mol – S^o(298.15 K) = 47.605 J/mol-K – p^o = 0.1 MPa









Zirconium, Ion (Zr^+) – Zr1+(g)

MW = 91.22 g/mol – $\Delta_f H^o$(298.15 K) = 1277.88 kJ/mol – S^o(298.15 K) = 183.639 J/mol-K – p^o = 0.1 MPa

