CALCULATED INTERACTION ENERGY
AND DIPOLE MOMENT IN COLLISIONS
OF TWO HYDROGEN MOLECULES

by R. W. Patch

Lewis Research Center
Cleveland, Ohio

NATIONAL AERONAUTICS AND SPACE ADMINISTRATION • WASHINGTON, D. C. • OCTOBER 1969
Calculations were carried out using orthogonalized valence-bond theory. Methods of calculating accurate values of integrals are given. Linear, planar perpendicular, quadrilateral, and nonplanar perpendicular configurations were included with intermolecular distances from 2.5 to 5.5 bohrs (0.13 to 0.29 nm) and two internuclear distances. The interaction energies obtained appear to be valid for intermolecular distances less than 4.5 bohr (0.24 nm) but are too large at appreciably greater intermolecular distances due to insufficient allowance for electron correlation in the wave function. A simple average interaction energy (Morse potential) for practical applications is cited which correlates calculations in this report and experiments.
# CONTENTS

<table>
<thead>
<tr>
<th>Section</th>
<th>Page</th>
</tr>
</thead>
<tbody>
<tr>
<td>SUMMARY</td>
<td>1</td>
</tr>
<tr>
<td>INTRODUCTION</td>
<td>2</td>
</tr>
<tr>
<td>ANALYSIS</td>
<td>4</td>
</tr>
<tr>
<td>Model and Configurations</td>
<td>4</td>
</tr>
<tr>
<td>Energy and Interaction Energy</td>
<td>6</td>
</tr>
<tr>
<td>Löwdin transformation</td>
<td>6</td>
</tr>
<tr>
<td>Formal orthogonalized valence-bond structures</td>
<td>7</td>
</tr>
<tr>
<td>Matrix elements of the Hamiltonian</td>
<td>9</td>
</tr>
<tr>
<td>Energies</td>
<td>11</td>
</tr>
<tr>
<td>Average Interaction Energy</td>
<td>11</td>
</tr>
<tr>
<td>Electric Dipole Moment</td>
<td>13</td>
</tr>
<tr>
<td>Computer Programs</td>
<td>14</td>
</tr>
<tr>
<td>RESULTS AND DISCUSSION</td>
<td>15</td>
</tr>
<tr>
<td>Covalent-Ionic Valence-Bond Model of ( \text{H}_2 )</td>
<td>15</td>
</tr>
<tr>
<td>( \text{H}_2 - \text{H}_2 ) Energy and Interaction Energy</td>
<td>15</td>
</tr>
<tr>
<td>( \text{H}_2 - \text{H}_2 ) Average Interaction Energy</td>
<td>19</td>
</tr>
<tr>
<td>( \text{H}_2 - \text{H}_2 ) Electric Dipole Moment and Its Derivatives</td>
<td>19</td>
</tr>
<tr>
<td>Comparison of ( \text{H}_2 - \text{H}_2 ) Results with Other Investigators</td>
<td>20</td>
</tr>
<tr>
<td>Interaction energy</td>
<td>20</td>
</tr>
<tr>
<td>Average interaction energy</td>
<td>22</td>
</tr>
<tr>
<td>Electric dipole moment</td>
<td>23</td>
</tr>
<tr>
<td>CONCLUDING REMARKS</td>
<td>23</td>
</tr>
<tr>
<td>APPENDIXES</td>
<td></td>
</tr>
<tr>
<td>A - SYMBOLS</td>
<td>25</td>
</tr>
<tr>
<td>B - EVALUATION OF KINETIC ENERGY INTEGRALS</td>
<td>28</td>
</tr>
<tr>
<td>C - EVALUATION OF NUCLEAR ATTRACTION INTEGRALS</td>
<td>30</td>
</tr>
<tr>
<td>D - EVALUATION OF TWO-ELECTRON INTEGRALS</td>
<td>33</td>
</tr>
<tr>
<td>REFERENCES</td>
<td>39</td>
</tr>
</tbody>
</table>
CALCULATED INTERACTION ENERGY AND DIPOLE MOMENT
IN COLLISIONS OF TWO HYDROGEN MOLECULES

by R. W. Patch

Lewis Research Center

SUMMARY

Ab initio calculations of the interaction energy and electric dipole moment of two colliding hydrogen molecules were carried out on a digital computer using the orthogonalized valence-bond theory of McWeeny. Methods of calculating accurate values of all molecular integrals are given in some detail for cases with equal or unequal orbital exponents (screening constants). Linear, planar perpendicular, quadrilateral, and non-planar perpendicular configurations were included with intermolecular distances from 2.5 to 5.5 bohrs (0.13 to 0.29 nm) and internuclear distances in the hydrogen molecules of 1.401446 and 1.450000 bohrs (0.0741599 and 0.0767292 nm). A weighted average of the interaction energies of the four equilibrium configurations was then taken for each intermolecular distance.

The interaction energies obtained appear to be valid for intermolecular distances less than 4.5 bohrs (0.24 nm) but are too large at appreciably greater intermolecular distances because of insufficient allowance for electron correlation in the wave function. A simple average interaction energy (Morse potential) for practical applications is cited which correlates experiments for intermolecular distances greater than 4.0 bohrs (0.21 nm) and correlates the calculations in this report for intermolecular distances less than 4.0 bohrs (0.21 nm).

The electric dipole moments obtained with the full valence-bond model appear to be approximately correct for intermolecular distances up to 4.5 bohrs (0.24 nm), but for appreciably larger distances are too small because of limitations of the model. Derivatives of the electric dipole moment with respect to the internuclear distances were also calculated.
INTRODUCTION

In high-temperature propulsion devices such as gas-core nuclear rockets, an important mechanism of heat transfer is radiant energy exchange between volumes of gas and between the gas and the wall (refs. 1 and 2). When such devices use high pressure hydrogen, it is necessary to know the strength of pressure-induced infrared absorption to calculate the heat transfer. This strength depends on the interaction energy and electric dipole moment of two colliding hydrogen molecules (H₂ - H₂) and has not been measured much above room temperature. At the temperatures that occur in these high-temperature propulsion devices, smaller intermolecular distances become important in calculating the strength.

The object of the work herein described was to calculate interaction energies and electric dipole moments for small intermolecular distances. The results should also be useful for calculating high-temperature transport properties of gases containing hydrogen and for calculating the equation of state of high-pressure hydrogen for application to light-gas guns.

Past work on H₂ - H₂ interaction energy may be divided into ab initio calculations, semiempirical calculations, and experiments. These three types of work will be discussed in subsequent paragraphs. In none of these interaction energy calculations was the dipole moment calculated, but the dipole moment by itself was calculated in a semiempirical calculation by another investigator.

Ab initio calculations for H₂ - H₂ have been carried out in 10 studies. For those unfamiliar with this type of calculation, a review of the methods used has been given by Slater (ref. 3). The H₂ - H₂ calculations may be divided into two groups: those where approximations were used for some molecular integrals, and those where all molecular integrals were calculated accurately. Since the interaction energy is a sensitive function of the molecular integrals, this is an important distinction. Griffing and Vanderslice (ref. 4) calculated energy of linear H₄⁺. Griffing and Maček (ref. 5) calculated the energy of square H₂⁺. Ruffa and Griffing (ref. 6) computed the interaction energy of linear H₂ - H₂. All three of these studies used approximations for some of the molecular integrals. Accurate values of the integrals were used by Taylor (ref. 7) for the energy of linear H₄⁻, Magnasco and Musso (refs. 8 to 12) for the energy of quadrilateral and skewed H₂ - H₂, Magnasco, Musso, and McWeeny (ref. 13) for the energy of quadrilateral and skewed H₂ - H₂, and by Schwartz and Schaad (ref. 14) for the energy of linear H₄⁻. None of the studies using accurate values of the integrals were done for a sufficient number of configurations to permit proper averaging of the interaction energies over nuclear configuration space, although this is necessary for practical applications.

Semiempirical calculations of H₂ - H₂ interaction energy have been made by
Figure 1. - Nine configurations of two hydrogen molecules used in calculations. Intermolecular distance \( R_t \) had values of 2.5, 3.5, 4.5, and 5.5 bohrs (1 bohr = 5.29167 \times 10^{-11} \text{ m}). Intennuclear distances in the figure are in bohrs (not to scale).
Margenau (ref. 15) (corrected by Evett and Margenau in ref. 16), Mason and Hirschfelder (ref. 17), Vanderslice and Mason (ref. 18), and Abrams, Patel, and Ellison (ref. 19). A semiempirical calculation of the $\text{H}_2 - \text{H}_2$ electric dipole moment has been made by Van Kranendonk and Kiss (ref. 20).

Experiments pertinent to $\text{H}_2 - \text{H}_2$ average interaction energy involve hydrogen viscosity and equation of state measurements. A large number of measurements of these two types have been made, but most were not at high enough temperatures to give any indication of the $\text{H}_2 - \text{H}_2$ interaction energy at small intermolecular distances. However, Guevara and Wageman (ref. 21) measured hydrogen viscosity to 2340 K. Fisher (ref. 22) fit Guevara and Wageman’s data together with lower temperature viscosity and equation-of-state data with a Morse intermolecular potential, which should therefore be valid down to an intermolecular distance of 4.0 bohrs (0.21 nm).

In this study a full valence-bond, ab initio calculation of $\text{H}_2 - \text{H}_2$ interaction energy and electric dipole moment was carried out with accurate values for all molecular integrals. In order to cover nuclear configuration space, linear, planar perpendicular, quadrilateral, and nonplanar perpendicular configurations (see fig. 1) were included with intermolecular distances from 2.5 to 5.5 bohr (0.13 to 0.29 nm). In four of the configurations (fig. 1(a), (c), (f), and (h)) both $\text{H}_2$ molecules had equilibrium internuclear distances. The average $\text{H}_2 - \text{H}_2$ interaction energy was then calculated for a given intermolecular distance by taking weighted averages of the interaction energies of these four configurations. In five of the configurations (fig. 1(b), (d), (e), (g), and (i)) one of the $\text{H}_2$ molecules had a nonequilibrium internuclear distance so that the derivatives of the dipole moment with respect to internuclear distance could be calculated from the electric dipole moments of all nine configurations. The calculations in this report differ from those of references 9 and 10 in the configurations calculated and in mathematical details.

The author is indebted to Professor V. Magnasco of the Istituto di Chimica Industriale dell’Università, Genoa, for providing listings and a deck of various computer programs for evaluating molecular integrals. Although these programs were not used for this report, they were useful for checking and provided ideas for the programs that were used.

ANALYSIS

Model and Configurations

The interaction of two $\text{H}_2$ molecules may be calculated by the orthogonalized valence-bond theory of McWeeny (refs. 23 and 24), which utilizes symmetrically orthogonalized atomic orbitals in the Born-Oppenheimer approximation. In this method
orthogonalized orbitals are formed from linear combinations of atomic orbitals. These are multiplied by spin eigenfunctions to give orthogonalized spin-orbitals. Formal orthogonalized valence-bond structures are formed from linear combinations of anti-symmetrized products of the orthogonalized spin-orbitals. The system wave function consists of linear combinations of these structures. This involved procedure merely amounts to taking the proper linear combination of the octuple products of four atomic orbitals and four spin eigenfunctions. This method has the advantage that it allows full configuration interaction without serious complexity. It has previously been applied to the interaction of two H$_2$ molecules by Magnasco and Musso (refs. 8 to 10).

The simplest possible model for a full valence-bond calculation was chosen. The two H$_2$ molecules were assumed to be in the ground electronic state so that for infinite separation each of the H$_2$ molecules could be described by the covalent-ionic valence-bond model of Weinbaum (ref. 25). In Weinbaum's model, two 1s atomic orbitals with equal orbital exponents (screening constants) were used. To be consistent, in the full valence-bond model used for H$_2$ - H$_2$ a set of four 1s atomic orbitals $\chi_p$ centered on the protons was selected. The respective orbital exponents $\xi_p$ were assumed to be the same as in the respective isolated H$_2$ molecules described by the Weinbaum model. If both molecules had the same internuclear distances, all orbital exponents were equal. If the two molecules had different internuclear distances, two orbital exponents had one value, and the other two had a different value. The four normalized 1s atomic orbitals $\chi_p$ were

$$\chi_p(i) = \left(\frac{\xi_p^3}{\pi}\right)^{1/2} e^{-\xi_p r_{pi}} \quad p = a, b, c, d$$

where the $i$ in parentheses indicates the $i^{th}$ electron, and $r_{pi}$ is the distance between proton $p$ and electron $i$ (symbols are given in appendix A). An orbital $\chi_p(i)$, of course, reduces to the wave function for the ground state of an H atom if $\xi_p = 1.0$ bohr$^{-1}$.

Four equilibrium configurations (fig. 1(a), (c), (f), and (h)) were chosen so that interaction energy could be conveniently averaged over molecular orientations. The experimental equilibrium internuclear distance of 1.401446 bohrs (0.0741599 nm) for H$_2$ (refs. 26 and 27) was used rather than Weinbaum's value of 1.4166 bohrs (0.074962 nm) (ref. 25) because of the ultimate application of this report to pressure induced absorption.

The other five configurations in figure 1 possessed a nonequilibrium internuclear distance so the derivatives of the electric dipole moment with respect to internuclear distance could be calculated from the nine configurations.

In all nine configurations the intermolecular distance $R_r$ was varied. It had values of 2.5, 3.5, 4.5, and 5.5 bohrs (1 bohr = 0.0529167 nm).
Energy and Interaction Energy

To find the electronic energy of the H$_2$ - H$_2$ complex, a Löwdin transformation was performed on the orbitals, resulting in orthogonalized orbitals. Formal orthogonalized valence-bond structures were formed, and the matrix elements of the Hamiltonian were found. An eigenvalue problem was then solved to obtain the electronic energy. The nuclear repulsion energy was added to this to get the energy of the complex. The energies of the two infinitely separated H$_2$ molecules were subtracted from the energy of the complex to yield the interaction energy. These steps are given in the following paragraphs.

Löwdin transformation. - The symmetrical orthogonalization procedure of Löwdin (ref. 28) was applied to the $\chi_p$. The overlap integral is

$$S_{pq} = \int \chi_p^{(1)} \chi_q^{(1)} dv_1$$

Let

$$w_{pq} = \frac{\xi_p + \xi_q}{2} R_{pq}$$

where $R_{pq}$ is the distance between protons $p$ and $q$. If $\xi_p = \xi_q$, $S_{pq}$ is given by Slater (ref. 3, p. 50). If $\xi_p \neq \xi_q$,

$$S_{pq} = \frac{\sqrt{\xi_p^3 \xi_q^3}}{4} \frac{R_{pq}^3}{4} \left[ \frac{e^{\tilde{q}} - e^{-\tilde{q}}}{\tilde{q}} \cdot \frac{2}{w_{pq}^2} \left( 1 + \frac{w_{pq}^2}{2} \right) + e^{-\tilde{q}} \left( \frac{1}{\tilde{q}} + \frac{2}{\tilde{q}^2} + \frac{2}{\tilde{q}^3} \right) \right] \cdot e^{-w_{pq}}$$

where

$$\tilde{q} = \frac{\xi_p - \xi_q}{2} R_{pq}$$
It is desired to find the Löwdin transformation matrix \( M^{-1/2}_\chi \) such that

\[
\chi' = \chi M^{-1/2}_\chi
\]  

(6)

where \( \chi \) is a row vector with elements \( \chi_p \), and \( \chi' \) is a row vector of the four orthogonalized orbitals \( \chi'_p \). First the eigenvalue problem

\[
\widetilde{U}S\widetilde{U} = d
\]

(7)

is solved, where \( S \) is a matrix with elements \( S_{pq} \). The matrix \( d \) is diagonal, with eigenvalues as diagonal elements. Thus

\[
\begin{bmatrix}
d^{-1/2}_{11} & 0 & 0 & 0 \\
0 & d^{-1/2}_{22} & 0 & 0 \\
0 & 0 & d^{-1/2}_{33} & 0 \\
0 & 0 & 0 & d^{-1/2}_{44}
\end{bmatrix}
\]

(8)

Finally,

\[
M^{-1/2}_\chi = \widetilde{U}d^{-1/2}_\chi \widetilde{U}
\]

(9)

Use of equation (6) then accomplishes the Löwdin transformation to obtain \( \chi' \).

Formal orthogonalized valence-bond structures. - To produce the proper antisymmetry, the \( \chi'_p \) are combined in antisymmetrized products of orthogonalized spin-orbitals.

\[
(pqrs) = (4!)^{-1/2} \det \left[ \chi'_p(1)\alpha(1), \chi'_q(2)\alpha(2), \chi'_r(3)\beta(3), \chi'_s(4)\beta(4) \right]
\]

(10)

where \( \alpha(1) \) is the spin eigenfunction of electron 1 with the component of spin angular momentum along the axis of quantization equal to \( \hbar/2 \), and \( \beta(3) \) is the spin eigenfunction of electron 3 with the component of spin angular momentum along the axis of quantization equal to \( -\hbar/2 \). The shorthand notation is given on the left side of equation (10), wherein a bar indicates \( \beta \) spin and no bar indicates \( \alpha \) spin. To facilitate computing for some purposes, the \( (pqrs) \) are arranged in dictionary order in a \( 6 \times 6 \) matrix (ref. 29).
For our purposes we will regard the $\mathbf{pqFS}$ as a 36-element row vector with the elements numbered down the first column in equation (11), then down the second column, etc. When numbered in this way, set

$$\varphi_i = (pqFS) \quad i = 1, 2, \ldots, 36 \quad (12)$$

where $p, q, f, F$, and $s$ are given in equation (11).

The formal orthogonalized valence-bond structures $\psi_j$ are linear combinations of the $\varphi_i$. These $\psi_j$ are eigenfunctions of the total electron spin with total spin angular momentum of 0 and component of the total spin angular momentum along the axis of quantization of 0. They have been given by Magnasco and Musso (ref. 9) and in the present nomenclature are

\[
\begin{align*}
\psi_1 &= \frac{1}{2} (-\varphi_{11} - \varphi_{16} - \varphi_{21} - \varphi_{26}) \\
\psi_2 &= \frac{1}{2} (\varphi_5 - \varphi_{16} - \varphi_{21} + \varphi_{31}) \\
\psi_3 &= 2^{-1/2} (-\varphi_{10} - \varphi_{20}) \\
\psi_4 &= 2^{-1/2} (\varphi_{17} - \varphi_{27}) \\
\psi_5 &= 2^{-1/2} (-\varphi_9 - \varphi_{14}) \\
\psi_6 &= 2^{-1/2} (-\varphi_{23} - \varphi_{28}) \\
\psi_7 &= 2^{-1/2} (-\varphi_{4} - \varphi_{19}) \\
\psi_8 &= 2^{-1/2} (\varphi_{18} + \varphi_{33}) \\
\psi_9 &= 2^{-1/2} (\varphi_{3} + \varphi_{13}) \\
\psi_{10} &= 2^{-1/2} (-\varphi_{24} - \varphi_{34}) \\
\psi_{11} &= 2^{-1/2} (\varphi_{5} + \varphi_{25}) \\
\psi_{12} &= 2^{-1/2} (\varphi_{12} + \varphi_{32}) \\
\psi_{13} &= 2^{-1/2} (-\varphi_{2} - \varphi_{7}) \\
\psi_{14} &= 2^{-1/2} (-\varphi_{30} - \varphi_{35}) \\
\psi_{15} &= -\varphi_{8} \\
\psi_{16} &= -\varphi_{15} \\
\psi_{17} &= -\varphi_{22} \\
\psi_{18} &= -\varphi_{29} \\
\psi_{19} &= \varphi_{1} \\
\psi_{20} &= \varphi_{36}
\end{align*}
\]
where formal singly polar and doubly polar structures are included. All $\psi_1$ are orthonormal (see ref. 30) except $\psi_1$ and $\psi_2$. These may be made orthonormal by the Schmidt orthogonalization procedure, which results in replacing $\psi_2$ in equation (13) by a new $\psi_2$ such that

$$
\psi_2 = \frac{\varphi_6}{3^{1/2}} + \frac{\varphi_11}{2 \cdot 3^{1/2}} - \frac{\varphi_{16}}{2 \cdot 3^{1/2}} - \frac{\varphi_{21}}{2 \cdot 3^{1/2}} + \frac{\varphi_{26}}{2 \cdot 3^{1/2}} + \frac{\varphi_{31}}{3^{1/2}}
$$

Equations (13) (excluding $\psi_2$) and (14) can be represented by

$$\psi = \varphi V$$

where $V$ is a 36 by 20 matrix. The wave function of the system was taken as

$$\Psi = \psi C$$

where $C$ is a column vector determined by minimizing the electronic energy $E_e$. This is accomplished by the variation principle for orthonormal functions (ref. 31) which poses an eigenvalue problem

$$\hat{H}C = E_e C$$

where the smallest of the several eigenvalues $E_e$ is the ground state electronic energy, and $\hat{H}$ is a matrix whose elements are given in the following sections.

**Matrix elements of the Hamiltonian.** - The Hamiltonian operator for the electronic energy in atomic units with the hartree as the unit of energy is

$$\hat{H} = \sum_{i=1}^{4} \frac{-1}{2} \nabla_i^2 + \sum_{i=1}^{4} \sum_{p=a}^{d} \frac{-1}{r_{pi}} + \sum_{i=2}^{4} \sum_{j=1}^{i-1} \frac{1}{r_{ij}}$$

Here $\nabla_i^2$ is the Laplacian operator for the coordinates of electron $i$, and the $p$ summation is over protons $a$, $b$, $c$, and $d$. The elements of the $\hat{H}$ matrix in equation (17) are then

$$H_{nk} = \int \psi_n^* \hat{H} \psi_k d\tau$$

1
where \( \tau \) includes electron configuration and spin coordinates. From equations (15) and (19)

\[
\mathbf{H} = \tilde{V} \mathbf{H} \varphi V
\]

(20)

where

\[
(H_\varphi)_{ij} = \int \varphi_i \tilde{H} \varphi_j d\tau
\]

(21)

The matrix elements \((H_\varphi)_{ij}\) can be found from Slater's rules (ref. 30 after correction of a typographical error) if two kinds of integrals are known

\[
\bar{F}_{ij} = \int \int \chi_i'(1) \left( \frac{1}{2} \frac{\nabla^2 \chi_j(1)}{\chi_j(1)} + \sum_{p=a}^{d} \frac{1}{r_{p1}} \right) \chi_j'(1) d\nu_1
\]

(22)

\[
\bar{G}_{ijkl} = \int \int \int \chi_i'(1) \chi_j'(1) \chi_k'(2) \chi_l'(2) \frac{1}{r_{12}} \frac{1}{d\nu_1} d\nu_2
\]

(23)

where \( r_{12} \) is the distance from electron 1 to electron 2, \( d\nu_1 \) is the volume element for electron 1, and \( d\nu_2 \) is the volume element for electron 2. From equations (6), (22), and (23)

\[
\bar{F}_{ij} = \sum_{k=1}^{4} \left( M_{\chi}^{-1/2} \right)_{ki} \sum_{l=1}^{4} \left( M_{\chi}^{-1/2} \right)_{lj} \left( K_{kl} + \sum_{p=a}^{d} L_{klp} \right)
\]

(24)

\[
\bar{G}_{ijkl} = \sum_{m=1}^{4} \left( M_{\chi}^{-1/2} \right)_{mi} \sum_{n=1}^{4} \left( M_{\chi}^{-1/2} \right)_{nj} \sum_{o=1}^{4} \left( M_{\chi}^{-1/2} \right)_{ok} \sum_{p=1}^{4} \left( M_{\chi}^{-1/2} \right)_{pl} G_{mnop}
\]

(25)
where

\[ K_{kl} = \int x_k(1) \left( -\frac{1}{2} \nabla_1^2 \right) x_l(1) \, dv_1 \]  \hspace{1cm} (26)

\[ L_{jkp} = \int x_j(1)x_k(1)\left(-\frac{1}{r_{pl}}\right) \, dv_1 \]  \hspace{1cm} (27)

\[ G_{mnop} = \int \int x_m(1)x_n(1)x_o(2)x_p(2) \frac{1}{r_{12}} \, dv_1 \, dv_2 \]  \hspace{1cm} (28)

The methods of evaluating \( K_{kl} \), \( L_{jkp} \), and \( G_{mnop} \) are given in appendixes B to D, respectively.

**Energies.** - The energy \( E \) is the electronic energy plus the nuclear repulsion energy.

\[ E = E_e + \frac{1}{R_{ab}} + \frac{1}{R_{ac}} + \frac{1}{R_{ad}} + \frac{1}{R_{bc}} + \frac{1}{R_{bd}} + \frac{1}{R_{cd}} \]  \hspace{1cm} (29)

The interaction energy \( E_{\text{int}} \) is the energy \( E \) minus the energies \( E_m(r_1) \) and \( E_m(r_2) \) of the isolated \( \text{H}_2 \) molecules with internuclear distances \( r_1 \) and \( r_2 \), respectively.

\[ E_{\text{int}} = E - E_m(r_1) - E_m(r_2) \]  \hspace{1cm} (30)

The diatomic internuclear distances used in equation (30) were the same as in the \( \text{H}_2 - \text{H}_2 \) complex (see fig. 1). The \( \text{H}_2 \) energies \( E_m(r) \) were calculated for the covalent-ionic valence-bond model (ref. 25).

**Average Interaction Energy**

The average interaction energy \( \overline{E}_{\text{int}} \) of two \( \text{H}_2 \) molecules is, to a first approximation, the interaction energy for equilibrium diatomic internuclear distances averaged over all possible orientations of the two \( \text{H}_2 \) molecules. Since only four configurations
with equilibrium diatomic internuclear distances were considered in this report, an additional approximation was necessary. Instead of using all possible orientations, the nine orientations in figure 2 were used. These are equally probable if differences in interaction energy are neglected, and each is equivalent to configurations A, C, F, or H (see fig. 1). Hence,

$$
\overline{E}_{\text{int}} = \frac{1}{9} E_{\text{int}, a} + \frac{4}{9} E_{\text{int}, c} + \frac{2}{9} E_{\text{int}, f} + \frac{2}{9} E_{\text{int}, h}
$$

(31)

where the constants are weighting factors. A better value of \( \overline{E}_{\text{int}} \) can be found by including Boltzmann factors in the weighting (ref. 32, pp. 921-922), but then \( \overline{E}_{\text{int}} \) would be complicated by temperature dependence.
Electric Dipole Moment

Calculation of the dipole moment requires an electron population analysis in terms of the original nonorthogonal orbitals $\chi_p$. A column vector $C'$ may be defined by

$$\Psi = \phi C'$$  \hspace{1cm} (32)

From equations (15), (16), and (32)

$$C' = VC$$  \hspace{1cm} (33)

The electron population in the orbital density $\chi_p \chi_q$ irrespective of spin state is

$$P'_{mn} = \sum_{r=1}^{36} C'_r \sum_{s=1}^{36} C'_s \left(n_{rs}^{m\alpha n\alpha} + n_{rs}^{m\beta n\beta}\right)$$  \hspace{1cm} (34)

where $n_{rs}^{m\alpha n\alpha}$ and $n_{rs}^{m\beta n\beta}$ are coefficients of spin-orbital products and are called $n_{rs}^{\mu\nu}$ by Magnasco and Musso (refs. 8 and 29), who give rules for evaluating them (ref. 29).

The electron population in the orbital density $\chi_p \chi_q$ irrespective of spin state is (ref. 8)

$$P_{pq} = \sum_{m=1}^{4} \left(M^{-1/2}\right)_{pm} \sum_{n=1}^{4} P'_{mn} \left(M^{-1/2}\right)_{qn}$$  \hspace{1cm} (35)

The normalized 1-electron population is

$$P_{pq} = P_{pq} S_{pq}$$  \hspace{1cm} (36)

where the four diagonal elements of the matrix $P$ give the atomic populations and the off-diagonal elements give the overlap populations. For instance, the population in the overlap between orbitals centered on protons $a$ and $b$ is $P_{ab} + P_{ba}$. The elements of the $P$ matrix, of course, add to 4, the number of electrons.

To find the dipole moment from $P$, it is necessary to know locations of the atomic charge centers and the overlap charge centers. Cartesian coordinates with the origin half way between the molecules were chosen (see fig. 1(a)). The $j^{th}$ Cartesian coordinate of either type of charge center was designated $C_{pqj}$, where $p$ and $q$ correspond to
the two orbitals $\chi_p$ and $\chi_q$ involved in an overlap charge or where $p$ and $q$ are equal and correspond to the one orbital involved in an atomic charge. The $C_{pqj}$ for atomic charges are, of course, identical to the coordinates of the protons $C_{pj}$. For overlap charges

$$C_{pqj} = \left(\frac{1}{2} - \Delta_{pq}\right) C_{pj} + \left(\frac{1}{2} + \Delta_{pq}\right) C_{qj}$$ (37)

where $\Delta_{pq}$ is the displacement of the overlap charge center from the $p$-$q$ geometric center, measured in the direction from $p$ towards $q$ and expressed in units of $R_{pq}$. If $\zeta_p = \zeta_q$, $\Delta_{pq} = 0$ by symmetry. If $\zeta_p \neq \zeta_q$.

$$\Delta_{pq} = \frac{R_{pq}^3 e^{-w_{pq}}}{8s_{pq}^3 w_{pq}^2} \left\{ e^{-\tilde{q}} \left( \frac{1}{q} - \frac{1}{q^2} \right) - e^{-\tilde{q}} \left( \frac{1}{q^2} - \frac{1}{q^4} \right) \left( \frac{6}{w_{pq}} + \frac{6}{w_{pq}^2} + 3 + w_{pq} \right) \right. $$

$$+ \left[ e^{-\tilde{q}} \left( \frac{1}{q^2} + \frac{3}{q^4} + \frac{6}{q^6} + \frac{6}{q^8} \right) + e^{-\tilde{q}} \left( \frac{1}{q^4} + \frac{3}{q^6} + \frac{6}{q^8} + \frac{6}{q^{10}} \right) \right] (1 + w_{pq}) \right\} \right.$$ (38)

The $j$th component of the dipole moment in atomic units is then

$$\mu_j = - \sum_{p=1}^{4} \sum_{q=1}^{4} \rho_{pq} \overline{C}_{pqj}$$ (39)

By symmetry, only the $z$ component may be nonzero.

**Computer Programs**

Calculation of the energy, interaction energy, and electric dipole moment of $H_2 - H_2$ was programmed in FORTRAN IV, IBM version 13, for an IBM 7094 digital computer. One program calculated all the three- and four-center molecular integrals and had an average running time of 8.09 minutes per case. A second program did the remaining calculations and had an average running time of 0.38 minute per case.
RESULTS AND DISCUSSION

In this section results are given for the covalent-ionic valence-bond model of $H_2$. These results are then used to calculate $H_2 - H_2$ interaction energy and electric dipole moment. Comparisons are made with the results of other investigators.

Covalent-Ionic Valence-Bond Model of $H_2$

The treatment of the covalent-ionic valence-bond model of $H_2$ given by Weinbaum (ref. 25) was reformulated so values of $\zeta$ and $E_m$ could be calculated for given internuclear distance $r$. The computer program of Dellepiane, Ferro, and Magnasco (ref. 33) was used to calculate the necessary integral logarithms. The results are given in table I. For an internuclear distance of 1.4166 bohrs (0.074962 nm), we got

TABLE I. ENERGY AND ORBITAL EXPONENT FOR COVALENT-IONIC VALENCE-BOND MODEL OF HYDROGEN MOLECULE

<table>
<thead>
<tr>
<th>Internuclear distance, $r$ (bohrs)</th>
<th>Energy, $E_m$ (hartrees)</th>
<th>$\zeta$ (bohrs$^{-1}$)</th>
<th>Orbital exponent, $\zeta$ (m$^{-1}$)</th>
</tr>
</thead>
</table>
| 1.300000 | -1.14462 | 1.2240 | 1.197 bohrs$^{-1}$, but Weinbaum got $\zeta = 1.193$ bohrs$^{-1}$. However, our energy agrees exactly with Weinbaum's. The energy was not very sensitive to the value of $\zeta$. Values of $E_m$ and $\zeta$ from table I were used in the following $H_2 - H_2$ calculations.

$H_2 - H_2$ Energy and Interaction Energy

Energy and interaction energy of $H_2 - H_2$ were calculated from equations (29) and (30), respectively, and are given in table II for nine nuclear configurations and four intermolecular distances. For a given intermolecular distance, the linear configurations...
TABLE II. - ENERGY AND ELECTRIC DIPOLE MOMENT FOR TWO HYDROGEN MOLECULES

<table>
<thead>
<tr>
<th>Intermolecular distance, $R_r$ bohrs</th>
<th>Configuration $m$</th>
<th>Energy, $E$</th>
<th>Interaction energy, $E_{int}$</th>
<th>$z$-component of electric dipole moment $b$</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td></td>
<td>hartree</td>
<td>J</td>
<td>hartree</td>
</tr>
<tr>
<td>2.5</td>
<td>1.3229×10^{-10}</td>
<td>A</td>
<td>-2.04492</td>
<td>-8.9147×10^{-18}</td>
</tr>
<tr>
<td></td>
<td></td>
<td>B</td>
<td>-2.03489</td>
<td>-8.8710</td>
</tr>
<tr>
<td></td>
<td></td>
<td>C</td>
<td>-2.15035</td>
<td>-9.3743</td>
</tr>
<tr>
<td></td>
<td></td>
<td>D</td>
<td>-2.15018</td>
<td>-9.3736</td>
</tr>
<tr>
<td></td>
<td></td>
<td>E</td>
<td>-2.14658</td>
<td>-9.3579</td>
</tr>
<tr>
<td></td>
<td></td>
<td>F</td>
<td>-2.18153</td>
<td>-9.5102</td>
</tr>
<tr>
<td></td>
<td></td>
<td>G</td>
<td>-2.18060</td>
<td>-9.5062</td>
</tr>
<tr>
<td></td>
<td></td>
<td>H</td>
<td>-2.18434</td>
<td>-9.5225</td>
</tr>
<tr>
<td></td>
<td></td>
<td>I</td>
<td>-2.19410</td>
<td>-9.5650</td>
</tr>
<tr>
<td>3.5</td>
<td>1.8521×10^{-10}</td>
<td>A</td>
<td>-2.25667</td>
<td>-9.8378×10^{-18}</td>
</tr>
<tr>
<td></td>
<td></td>
<td>B</td>
<td>-2.25540</td>
<td>-9.8323</td>
</tr>
<tr>
<td></td>
<td></td>
<td>C</td>
<td>-2.26824</td>
<td>-9.8882</td>
</tr>
<tr>
<td></td>
<td></td>
<td>D</td>
<td>-2.26800</td>
<td>-9.8872</td>
</tr>
<tr>
<td></td>
<td></td>
<td>E</td>
<td>-2.26735</td>
<td>-9.8844</td>
</tr>
<tr>
<td></td>
<td></td>
<td>F</td>
<td>-2.27440</td>
<td>-9.9151</td>
</tr>
<tr>
<td></td>
<td></td>
<td>G</td>
<td>-2.27412</td>
<td>-9.9139</td>
</tr>
<tr>
<td></td>
<td></td>
<td>H</td>
<td>-2.27456</td>
<td>-9.9158</td>
</tr>
<tr>
<td></td>
<td></td>
<td>I</td>
<td>-2.27706</td>
<td>-9.9267</td>
</tr>
<tr>
<td>4.5</td>
<td>2.3813×10^{-10}</td>
<td>A</td>
<td>-2.28895</td>
<td>-9.9785×10^{-18}</td>
</tr>
<tr>
<td></td>
<td></td>
<td>B</td>
<td>-2.28874</td>
<td>-9.9776</td>
</tr>
<tr>
<td></td>
<td></td>
<td>C</td>
<td>-2.29084</td>
<td>-9.9868</td>
</tr>
<tr>
<td></td>
<td></td>
<td>D</td>
<td>-2.29082</td>
<td>-9.9867</td>
</tr>
<tr>
<td></td>
<td></td>
<td>E</td>
<td>-2.29070</td>
<td>-9.9861</td>
</tr>
<tr>
<td></td>
<td></td>
<td>F</td>
<td>-2.29194</td>
<td>-9.9916</td>
</tr>
<tr>
<td></td>
<td></td>
<td>G</td>
<td>-2.29193</td>
<td>-9.9915</td>
</tr>
<tr>
<td></td>
<td></td>
<td>H</td>
<td>-2.29193</td>
<td>-9.9915</td>
</tr>
<tr>
<td></td>
<td></td>
<td>I</td>
<td>-2.29257</td>
<td>-9.9943</td>
</tr>
<tr>
<td>5.5</td>
<td>2.9104×10^{-10}</td>
<td>A</td>
<td>-2.29453</td>
<td>-10.0028×10^{-18}</td>
</tr>
<tr>
<td></td>
<td></td>
<td>B</td>
<td>-2.29456</td>
<td>-10.0030</td>
</tr>
<tr>
<td></td>
<td></td>
<td>C</td>
<td>-2.29484</td>
<td>-10.0042</td>
</tr>
<tr>
<td></td>
<td></td>
<td>D</td>
<td>-2.29490</td>
<td>-10.0045</td>
</tr>
<tr>
<td></td>
<td></td>
<td>E</td>
<td>-2.29488</td>
<td>-10.0044</td>
</tr>
<tr>
<td></td>
<td></td>
<td>F</td>
<td>-2.29500</td>
<td>-10.0049</td>
</tr>
<tr>
<td></td>
<td></td>
<td>G</td>
<td>-2.29506</td>
<td>-10.0052</td>
</tr>
<tr>
<td></td>
<td></td>
<td>H</td>
<td>-2.29499</td>
<td>-10.0048</td>
</tr>
<tr>
<td></td>
<td></td>
<td>I</td>
<td>-2.29519</td>
<td>-10.0057</td>
</tr>
</tbody>
</table>

aSee fig. 1.

bSee fig. 1(a) for coordinates. Other components of the electric dipole moment are zero by symmetry.
(a) Linear configuration A (see fig. 1(a)).

(b) Planar perpendicular configuration C (see fig. 1(c)).

Figure 3. Interaction energy for four configurations of two hydrogen molecules, each configuration having equilibrium internuclear distances in both molecules.
This report (eq. (30)) - Mason and Hirschfelder (ref. 17)

(c) Quadrilateral configuration F (see fig. 1(f)). The curves for this report and from reference 9 coincide from $R_r = 2.5$ to 4.2478 bohrs.

(d) Nonplanar perpendicular configuration H (see fig. 1(h)).

Figure 3. - Concluded.
had the highest interaction energies, followed by the planar perpendicular configurations. The quadrilateral and nonplanar perpendicular configurations had the lowest interaction energies. It is not safe to draw any conclusions about the change in interaction energy between equilibrium and nonequilibrium configurations because the experimental equilibrium internuclear distance was used in this report rather than Weinbaum's value (ref. 25 and table I). Values of interaction energy in table II are plotted in figure 3 for the four equilibrium configurations. The interaction energies varied approximately exponentially with intermolecular distance.

H₂ - H₂ Average Interaction Energy

The average interaction energy of H₂ - H₂ was calculated from equation (31) for four intermolecular distances. Results are given in table III and are plotted in figure 4.

**TABLE III. - AVERAGE INTERACTION ENERGY FOR TWO HYDROGEN MOLECULES WITH EQUILIBRIUM INTERNUCLEAR DISTANCES**

<table>
<thead>
<tr>
<th>Intermolecular distance, ( R )</th>
<th>Average interaction energy, ( E_{\text{int}} )</th>
</tr>
</thead>
<tbody>
<tr>
<td>( \text{bohrs} )</td>
<td>( m )</td>
</tr>
<tr>
<td>2.5</td>
<td>1.3229\times10^{-10}</td>
</tr>
<tr>
<td>3.5</td>
<td>1.8521</td>
</tr>
<tr>
<td>4.5</td>
<td>2.3813</td>
</tr>
<tr>
<td>5.5</td>
<td>2.9104</td>
</tr>
</tbody>
</table>

\( ^a \) Equilibrium internuclear distances of both molecules were 1.401446 bohrs (7.4160\times10^{-11} m) (ref. 26).

H₂ - H₂ Electric Dipole Moment and Its Derivatives

The electric dipole moment of H₂ - H₂ was calculated from equation (39) for the nine configurations in figure 1 and four intermolecular distances. The \( x \) and \( y \) components were all essentially zero, as required by symmetry. The \( z \) component is given in table II. The planar perpendicular configurations had the largest magnitude of dipole...
moment because of their high degree of asymmetry. Equilibrium configurations A, F, and H had zero dipole moment by symmetry. The z component of dipole moment of equilibrium configuration C is plotted in figure 5. It varies approximately exponentially with internuclear distance. The derivatives of the dipole moment with respect to internuclear distances were calculated from the dipole moments of the nine configurations (see table II) and are shown in figure 6.

Comparison of H₂ - H₂ Results with Other Investigators

Interaction energy. - Comparisons of $E_{\text{int}}$ from equation (30) with ab initio results of other investigators are given in figure 3 for four H₂ - H₂ configurations with equilibrium internuclear distances in both molecules.
This report (eq. (39))

Large $R_{r}$ approximation using $Q_{a}$ from ref. 34 and $a$ from ref. 35

Large $R_{r}$ approximation as above plus overlap contribution from ref. 20

Figure 5. $z$ component of electric dipole moment for planar perpendicular configuration of two hydrogen molecules. Coordinates are shown in figure 1(a).

Figure 3(a) is for the linear configuration. The interaction energies of Griffing and Vanderslice (ref. 4) and Ruffa and Griffing (ref. 6) are high, possibly because of inaccurate molecular integrals. The energy of Taylor (ref. 7) is also high, possibly because he set all the orbital exponents equal to one. The energies of Schwartz and Schaad (ref. 14) are lower than our energies. The variation principle states that $E$ is an upper limit, but gives no information on $E_{\text{int}}$. Consequently, this principle cannot be used to show that Schwartz and Schaad's $E_{\text{int}}$ is best. However, since they used 24 orbitals compared with four in this report, their $E_{\text{int}}$ is probably best for the range of $R_{r}$ they covered. The theory of Mason and Hirschfelder (ref. 17) is believed best at large $R_{r}$ but is not accurate at small $R_{r}$ because it is principally a perturbation calculation. Its superiority at large $R_{r}$ is due to its accurate semiempirical treatment of dispersion energy.

Figures 3(b) to (d) are for the planar perpendicular, quadrilateral, and nonplanar perpendicular configurations, respectively. The $E_{\text{int}}$ of Mason and Hirschfelder
Planar perpendicular

Nonplanar perpendicular

Planar perpendicular

Average interaction energy. - Comparison of $E_{\text{int}}$ from equation (31) with results of other investigators are given in figure 4 for $H_2 - H_2$ with equilibrium internuclear distances in both molecules. The Lennard-Jones 6-12 potential (ref. 32) was fitted to data from moderate temperature equation-of-state experiments and hence is not accurate at small $R_r$. The Morse potential (ref. 22) was fitted to moderate temperature equation-of-state and viscosity data and also to viscosity data up to 2340 K (ref. 21); therefore, it should be valid down to $R_r = 4.0$ bohrs ($2.1 \times 10^{-10}$ m). However, it is obviously wrong for small enough $R_r$ because it does not approach infinity as $R_r$ approaches zero. Thus its agreement with equation (31) for $2.5 \leq R_r \leq 4$ bohrs ($0.13 \leq R_r \leq 0.21$ nm) is surprising. Two semiempirical theories (refs. 17 and 18) are included in figure 4 for reference only. I believe the excessive values of $E_{\text{int}}$ from equation (31) for large $R_r$
are due to insufficient allowance for electron correlation in the wave function. This causes the magnitude of the dispersion energy (which is a negative quantity) to be low.

**Electric dipole moment.** Comparisons of $\mu_z$ from equation (39) with results of other investigators are given in figure 5 for the planar perpendicular configuration. The large $R_T$ approximation was obtained by considering the polarization of each $H_2$ molecule by the electric quadrupole of the other. The scalar quadrupole moment $Q_S$ and polarizability $\alpha$ of $H_2$ were obtained from Kolos and Wolniewicz (refs. 34 and 35) and inspire confidence. The resulting $\mu_z$ varies as $1/R_T^4$ as shown and is correct without much doubt for large enough $R_T$. A curve is also given in figure 5 for the large $R_T$ approximation plus a semiempirical overlap contribution (ref. 20). The overlap contribution is not large, so the trend is essentially the same as for the large $R_T$ approximation alone. Equation (39) does not follow either dashed curve for $R_T > 4.5$ bohrs (0.24 nm) and, in fact, is even curved the wrong way. I believe this difficulty is due to the model used: at large $R_T$ it essentially describes two covalent-ionic valence-bond $H_2$ molecules, which have no polarizability perpendicular to their axes and probably have incorrect quadrupole moments. (The quadrupole moment is very sensitive to the electronic charge distribution.) Consequently, the orthogonalized valence-bond theory used in this report does not give reliable $\mu_z$ for $R_T > 4.5$ bohrs (0.24 nm) despite the inclusion of singly and doubly polar structures.

**CONCLUDING REMARKS**

Full valence-bond ab initio calculations of $H_2 - H_2$ interaction energy and electric dipole moment were carried out with accurate values for all molecular integrals. Linear, planar perpendicular, quadrilateral, and nonplanar perpendicular configurations were included with intermolecular distances from 2.5 to 5.5 bohr (0.13 to 0.29 nm). A weighted average of the interaction energies of the four equilibrium configurations was then taken.

The interaction energies obtained appear to be valid for intermolecular distances less than 4.5 bohr (0.24 nm) but are too large at appreciably greater intermolecular distances because of insufficient allowance for electron correlation in the wave function.

Fisher (ref. 22) has fitted a Morse potential to viscosity and equation-of-state data, and this potential should be valid down to intermolecular distances of 4.0 bohrs (0.21 nm), based on the temperatures at which the data were taken. The calculations in this report show that this Morse potential is also a good approximation for intermolecular distances between 2.5 and 4.0 bohrs (0.13 and 0.21 nm). Consequently, this Morse potential is useful for practical applications for all $R_T$ greater than 2.5 bohrs (0.13 nm).

The electric dipole moments obtained by the full valence-bond calculation appear to
be approximately correct for intermolecular distances up to 4.5 bohrs (0.24 nm). For appreciably larger distances they are too small due to limitations of the model. The derivatives of the dipole moment with respect to the internuclear distances of the \( \text{H}_2 \) molecules were also obtained. These are necessary to calculate pressure-induced vibrational absorption coefficients.

Lewis Research Center,
National Aeronautics and Space Administration,
Cleveland, Ohio, July 23, 1969,
122-28.
APPENDIX A

SYMBOLS

A  definite integral involving exponential
(abcd) shorthand notation for antisymmetrized product of orthogonalized spin-orbitals
C  column vector of coefficients of \( \psi_j \)
C' column vector of coefficients of \( \varphi_i \)
\( C_{pj} \)  \( j^{th} \) Cartesian coordinate of proton \( p \)
\( C_{ppj} \)  \( j^{th} \) Cartesian coordinate of center of atomic charge \( p \)
\( C_{pqj} \)  \( j^{th} \) Cartesian coordinate of center of overlap charge \( pq \)
\( C'_{r} \)  element of \( C' \)
\( d \)  matrix of overlap eigenvalues
\( d_{ij}^{-1/2} \)  inverse square root of \( d_{ij} \)
\( d_{ii}^{-1/2} \)  diagonal element of \( d_{ij}^{-1/2} \)
E  electronic and nuclear repulsion energy of complex
\( E_e \)  electronic energy
\( E_{\text{int}} \)  interaction energy
\( \overline{E}_{\text{int}} \)  average interaction energy
\( E_m \)  electronic and nuclear repulsion energy of diatomic molecule
\( F_{ij} \)  one-electron integral of orthogonalized orbitals
\( G_{ijkl} \)  two-electron integral of orthogonalized orbitals
\( G_i^k \)  definite integral involving exponential and Legendre function of first kind
\( G_{mnop} \)  two-electron integral of 1s orbitals
\( H \)  Hamiltonian matrix with basis functions \( \psi_j \)
\( \hat{H} \)  Hamiltonian operator for electrons
\( H_k \)  definite integral involving exponential and Legendre functions of first and second kinds
\( H_{nk} \)  element of \( H \)
\( H_{\varphi} \)  Hamiltonian matrix with basis functions \( \varphi_i \)
\( (H_{\varphi})_{ij} \)  element of \( H_{\varphi} \)
\( \hbar \)  Planck constant divided by \( 2\pi \)
J  center of orbital \( \chi_j \)
K  center of orbital \( \chi_k \)
\( K_{ij}^k \)  definite integral involving exponential and Legendre function of first kind
\( K_{kl} \)  kinetic energy integral
\( L_{jkp} \)  nuclear attraction integral
\( L_{ik}^j \)  definite integral involving exponential and Legendre function of second kind
M  center of orbital \( \chi_m \)
Löwdin transformation matrix

\[
\frac{1}{\chi} \frac{M^{-1/2}}{M^{-1/2}} \]

\( k \)

element of \( M^{-1/2} \)

\( N \)

center of orbital \( \chi_n \)

\( n_{\alpha \alpha}^{m} \)

coefficients of spin-orbital products with two \( \alpha \) spins

\( n_{\alpha \beta}^{m} \)

coefficients of spin-orbital products with two \( \beta \) spins

\( P \)

center of proton \( p \)

\( \mathcal{P} \)

matrix of \( \mathcal{P}_{pq} \)

\( P_k \)

Legendre function of first kind

\( P'_{mn} \)

electron population in the orbital density \( \chi'_m \chi'_n \) irrespective of spin state

\( P_{pq} \)

electron population in the orbital density \( \chi_p \chi_q \) irrespective of spin state

\( \mathcal{P}_{pq} \)

normalized one-electron population

\( P_k^j \)

constants in Legendre function of first kind

\( Q_k \)

Legendre function of second kind

\( Q_S \)

scalar quadrupole moment

\( \dot{q} \)

\( (\xi_p - \xi_q)R_{pq}/2 \) (subscripts vary)

\( \ddot{q} \)

\( (\xi_m - \xi_o)R_{mo} \)

\( \ddot{q} \)

\( (3\xi_m - \xi_p)R_{mp}/2 \)

\( R_{pq} \)

distance between points \( p \) and \( q \)

\( R_r \)

intermolecular distance (see fig. 1)

\( r, r_i \)

internuclear distance of diatomic molecule \( i \)

\( \hat{r} \)

\( (\xi_m + \xi_o)R_{mo} \)

\( \ddot{r} \)

\( (3\xi_m + \xi_p)R_{mp}/2 \)

\( r_{ij} \)

distance between electron or proton \( i \) and electron \( j \)

\( S \)

matrix of \( S_{pq} \)

\( S_{pq} \)

overlap integral for \( \chi_p \) and \( \chi_q \)

\( T \)

definite integral involving exponential

\( U \)

unitary matrix

\( \tilde{U} \)

transpose of \( U \)

\( U_{mn}(2) \)

electrostatic potential at electron \( 2 \) due to the bicentric charge distribution \( \chi_m(1)\chi_n(1) \)

\( u \)

variable of integration

\( V \)

matrix relating \( \bar{\psi} \) and \( \varphi \)

\( \tilde{V} \)

transpose of \( V \)

\( v_i \)

volume in configuration space for electron \( i \)

\( w, w_{pq} \)

\( (\xi_p + \xi_q)R_{pq}/2 \)

\( X_k, Z_k \)
\( x \) and \( z \) coordinates of point \( K \) in units of \( R_{pq}/2 \)

\( x, y, z \)

Cartesian coordinates

\( \alpha \)

polarizability

\( \alpha(i) \)

spin eigenfunction of electron \( i \) with component of spin angular momentum along the axis of quantization equal to \( \hbar/2 \)

\( \beta(i) \)

spin eigenfunction of electron \( i \) with component of spin angular momentum along the axis of quantization equal to \( -\hbar/2 \)
\( \Delta_{pq} \) displacement of overlap charge center for \( p-q \) geometric center, measured in direction from \( p \) towards \( q \) and expressed in units of \( R_{pq} \)

\( \zeta \) orbital exponent

\( \zeta_p \) orbital exponent of \( \chi_p \)

\( \theta, \mu, \nu \) spheroidal coordinates

\( \lambda \) \(-1/\mu\)

\( \mu_k \) \( \mu \) of point \( K \)

\( \mu_\infty \) finite upper limit for \( \mu \) integration

\( \mu(\langle) \) smaller of \( \mu_1 \) and \( \mu_2 \)

\( \mu(\rangle) \) larger of \( \mu_1 \) and \( \mu_2 \)

\( \tau \) volume in configuration and spin space of four electrons

\( \varphi \) row vector with elements \( \varphi_1 \)

\( \varphi_i \) antisymmetrized product of orthogonalized spin-orbitals

\( \chi \) row vector with elements \( \chi_p \)

\( \chi' \) row vector with elements \( \chi'_p \)

\( \chi_p(i) \) 1s orbital of electron \( i \) centered on proton \( p \)

\( \chi'_p(i) \) orthogonalized orbital of electron \( i \)

\( \Psi \) wave function of system

\( \psi \) row vector with elements \( \psi_j \)

\( \psi_j \) formal orthogonalized valence bond structure
APPENDIX B

EVALUATION OF KINETIC ENERGY INTEGRALS

Herein methods are given for evaluating kinetic energy integrals involving 1s atomic orbitals (eq. (1)) with equal or unequal orbital exponents. These integrals all have the form

\[ K_{kl} = \int x_k(1) \left( -\frac{1}{2} \frac{\nabla^2}{\nabla} \right) x_l(1) dv_1 \]

(B1)

Equal Orbital Exponents

Analytic expressions for \( K_{kl} \) for the two cases \( k = l \) and \( k \neq l \) have been given in reference 3 (p. 50) (but a typographical error must be corrected in the \( k \neq l \) case).

Unequal Orbital Exponents

If \( k \neq l \) and the orbital exponents are unequal, the integral in equation (B1) may be found analytically by the methods in reference 3.

\[ K_{kl} = -\frac{1}{8} R_{kl}^3 \sqrt{\frac{3}{k}} \left\{ \begin{array}{c} \xi_l \left[ e^{\tilde{q}} \left( \frac{1}{\tilde{q}} + \frac{2}{\tilde{q}^2} + \frac{2}{\tilde{q}^3} \right) + e^{-\tilde{q}} \left( \frac{1}{\tilde{q}} + \frac{2}{\tilde{q}^2} - \frac{2}{\tilde{q}^3} \right) \right] - 4 \frac{e^{\tilde{q}} - e^{-\tilde{q}}}{R_{kl}} \frac{1 + w_{kl}}{\tilde{q}} \frac{w_{kl}}{w_{kl}} \\
+ \frac{4}{R_{kl}} \left[ e^{-\tilde{q}} \left( \frac{1}{\tilde{q}} + \frac{1}{\tilde{q}^2} \right) + e^{\tilde{q}} \left( \frac{1}{\tilde{q}} - \frac{1}{\tilde{q}^2} \right) \right] e^{-w_{kl}} \end{array} \right\} \]

(B2)

where \( w_{kl} \) is given by equation (3) and where
\[ \tilde{q} = \frac{(\zeta_k - \zeta_l)R_{k\ell}}{2} \] (B3)

All \( K_{k\ell} \) were evaluated to within \( \pm 1 \) in the sixth significant digit.
APPENDIX C

EVALUATION OF NUCLEAR ATTRACTION INTEGRALS

Methods for evaluating nuclear attraction integrals involving 1s atomic orbitals (eq. (1)) with equal or unequal orbital exponents are given herein. These integrals all have the form

\[ L_{j kp} = \int x_j(1)x_k(1) \left( \frac{-1}{r_{p1}} \right) dv_1 \]  \hspace{1cm} (C1)

where \( r_{p1} \) is the distance from proton \( p \) to electron 1.

Equal Orbital Exponents

If all three subscripts of \( L_{j kp} \) are equal or if only two of them are equal, analytic expressions for \( L_{j kp} \) are given in reference 3 (p. 50). If all subscripts are different, numerical integration is required and is described in the following paragraphs.

Numerical integration of three-center integrals is best done in spheroidal coordinates (ref. 3). First, however, a set of Cartesian coordinates is selected with the origin half way between \( P \) and \( J \) and with the \( z \) axis in the direction from \( P \) to \( J \).

\[ 1 \]

30
The x axis is chosen so K is in the x-z plane with \( x_k \) zero or positive. The three spheroidal coordinates \( \theta, \mu, \) and \( \nu \) can now be defined. The azimuthal angle measured around the z axis as shown is called \( \theta \). The other two coordinates are

\[
\mu = \frac{r_{pl} + r_{jl}}{R_{pj}} \tag{C2}
\]

\[
\nu = \frac{r_{pl} - r_{jl}}{R_{pj}} \tag{C3}
\]

where \( r_{pl} \) is the distance from P to \( dv_1 \), and \( r_{jl} \) is the distance from J to \( dv_1 \). Surfaces of constant \( \mu \) are ellipsoids of revolution about the z axis with foci at P and J. Surfaces of constant \( \nu \) are hyperboloids of revolution about the z axis with foci at P and J except for \( \nu = 0 \), which is the x-y plane.

In these coordinates, equation (C1) becomes

\[
L_{jkp} = -\frac{R_{pj}^2}{2\pi} \sqrt{\xi_j \xi_k} \int_{-1}^{1} \left( R_{pj} \frac{(\mu - \nu)}{2} \int_{0}^{\pi} \exp \left[ -\frac{\xi_k}{2} \left( X_k^2 + Z_k^2 - 1 + \mu^2 + \nu^2 - 2\mu\nu Z_k - 2X_k \sqrt{\mu^2 - 1} \sqrt{1 - \nu^2 \cos \theta} \right)^{1/2} \right] d\theta d\nu d\mu \tag{C4}
\]

where \( X_k \) and \( Z_k \) are \( x_k \) and \( z_k \), respectively, in units of \( R_{pj}/2 \). The \( \theta \) integration was done by a single Gaussian quadrature. The \( \nu \) integration was done by using two Gaussian quadratures: one from -1 to \( \nu_k \) and one from \( \nu_k \) to 1. The \( \mu \) integration was also split into two parts. Integration from 1 to \( \mu_k \) was done by Gaussian quadrature. From \( \mu_k \) to an upper limit \( \mu_\infty \) the integration was done by transforming the integral so it was with respect to a variable \( \lambda = -1/\mu \) and applying Gaussian quadrature. (The author is indebted to Professor V. Magnasco, Universita di Genova, Genoa, for this idea.) In each case an appropriate finite \( \mu_\infty \) had to be calculated to replace \( \infty \) as the upper limit in the \( \mu \) integration.

Use of the coordinate system shown was superior to use of the coordinate system in reference 36 because it gave better accuracy for the same number of integration points for cases where J, K, and P were in line or almost in line. This was apparently because the integrand in equation (C4) is everywhere finite, which is not the case for the equation in different coordinates in reference 36.
Unequal Orbital Exponents

If \( j = k \neq p \), the only two orbital exponents that occur in the integral are identical. However, if \( p = j \neq k \) or \( p = k \neq j \), the unequal orbital exponents result in new expressions for \( L_{jkp} \). For the case \( p = k \neq j \)

\[
L_{jkp} = -\frac{1}{2} R_{jk}^2 \sqrt{\frac{3}{\tilde{\zeta}_j} \frac{3}{\tilde{\zeta}_k}} \left[ e^{\tilde{\zeta}_j} e^{-\tilde{\zeta}_k} \frac{1 + w_{jk}}{\tilde{\zeta}_j} - e^{-\tilde{\zeta}_j} \frac{1 + w_{jk}}{\tilde{\zeta}_k} - e^{\tilde{\zeta}_k} \frac{1 - \frac{1}{\tilde{\zeta}_j}}{\tilde{\zeta}_j^2} \right] e^{-w_{jk}} \]  

(C5)

where the methods of reference 3 were used and where

\[
\tilde{\zeta} = \frac{(\tilde{\zeta}_j - \tilde{\zeta}_k)R_{jk}}{2} \]  

(C6)

For the case where none of the subscripts of \( L_{jkp} \) are equal, equation (C4) is applicable.

All \( L_{jkp} \) were evaluated to within \( \pm 3 \) in the sixth significant digit.
APPENDIX D

EVALUATION OF TWO-ELECTRON INTEGRALS

Methods are given for evaluating two-electron integrals involving 1s atomic orbitals (eq. (1)) with equal or unequal orbital exponents. These integrals all have the form

\[ G_{mnop} = \int \int \chi_m(1)\chi_n(2)\chi_o(2)\chi_p(2) \frac{1}{r_{12}} dv_1 dv_2 \]  \hspace{1cm} (D1)

where \( r_{12} \) is the distance between the two electrons.

Equal Orbital Exponents

One-center integral. - An expression for \( G_{mnop} \) when all the subscripts are equal is given in reference 3 (p. 65).

Two-center Coulomb integral. - If \( m = n \) and \( o = p \), an expression for \( G_{mnop} \) is given in reference 3 (p. 50).

Two-center exchange integral. - If \( m = o \) and \( n = p \), an expression for \( G_{mnop} \) is given in reference 3 (p. 50), but it requires correction of a typographical error. The expression involves the integral logarithm, which was evaluated using the program of Dellepiane, Ferro, and Magnasco (ref. 33).

Two-center hybrid integral. - An expression for \( G_{mnop} \) if three of the subscripts are identical is given in reference 3 (p. 65).

Three-center Coulomb integral. - If \( m = n \neq o \neq p \) or \( m \neq n \neq o = p \), the integral must be integrated numerically. The method of Magnasco and Dellepiane (ref. 36) was used except that on the outer \( \mu \) integration range the \( \mu \) integral was transformed so that it was with respect to a variable \( \lambda = -1/\mu \) and Gaussian quadrature was applied.

Three-center exchange integral. - If one of the first two subscripts equalled one of the last two subscripts of \( G_{mnop} \), the integral was integrated numerically by the method of Magnasco and Dellepiane (ref. 37) except that on the outer \( \mu \) integration range the method of the preceding paragraph was used.

Four-center integral. - For cases where all the subscripts of \( G_{mnop} \) were different, an evaluation method has been given by Magnasco and Dellepiane (ref. 37). This method was used with three modifications to improve the accuracy: (1) on the outer \( \mu \) integration range the \( \mu \) integral was transformed so it was with respect to a variable \( \lambda = -1/\mu \), and Gaussian quadrature was applied; (2) in some cases three \( \mu \) integration
ranges were used instead of two; and (3) in some cases three \( \nu \) integration ranges were used instead of two.

### Unequal Orbital Exponents

**Two-center Coulomb integral.** - Here \( m = n \) and \( o = p \). Using the methods in reference 3 gives

\[
G_{mop} = \frac{1}{2} \frac{\xi^3_o R^2}{\xi^3_o R_{mo}} \left\{ \frac{2}{\xi^3_o R^3_{mo}} - 2e^{-2\xi_o R_{mo}} \left( \frac{1}{\xi^3_o R^3_{mo}} + \frac{1}{\xi^2_o R^2_{mo}} \right) - \frac{e^{-\hat{q}} - e^{-\hat{q}}}{q} \frac{e^{-\hat{r}}}{r^2} (1 + \hat{r}) \right. \\
\left. - \left[ e^{-\hat{q}} \left( \frac{1}{\hat{q}} + \frac{1}{\hat{q}^2} \right) + e^{-\hat{q}} \left( \frac{1}{\hat{q}^2} \right) \right] \frac{e^{-\hat{r}}}{\hat{r}} - \xi_m R_{mo} \frac{e^{-\hat{q}}}{\hat{q}} \frac{e^{-\hat{r}}}{\hat{r}^3} \left( 1 + \hat{r} + \frac{\hat{r}^2}{2} \right) \\
- \frac{1}{2} \xi_m R_{mo} \left[ e^{-\hat{q}} \left( \frac{1}{\hat{q}} + \frac{2}{\hat{q}^2} \right) + e^{-\hat{q}} \left( \frac{1}{\hat{q}^2} \right) \right] \frac{e^{-\hat{r}}}{\hat{r}} \right\} \tag{D2}
\]

where

\[
\hat{r} = (\xi_m + \xi_o) R_{mo} \tag{D3}
\]

and

\[
\hat{q} = (\xi_m - \xi_o) R_{mo} \tag{D4}
\]

**Two-center exchange integral.** - Here \( m = o \) and \( n = p \). This can be expressed as an infinite series. Spheroidal coordinates were used for both electrons. The quantity \( 1/r_{12} \) was expanded in the Newmann expansion (ref. 3, p. 266) with the result

\[
G_{mop} = \frac{1}{8} R^5 m^3 n^3 \sum_{k=0}^{\infty} (2k + 1) \left[ G^0_k(q) H_k(2, 2, w_{mn}) - 2G^0_k(q) G^2_k(q) H_k(0, 2, w_{mn}) + G^2_k(q) H_k(0, 0, w_{mn}) \right] \tag{D5}
\]
where \( w_{mn} \) is given by equation (3) and

\[
\tilde{q} = \frac{\zeta_m - \zeta_n}{2} \quad \text{(D6)}
\]

\[
G_k^i(\tilde{q}) = \int_{-1}^{1} \nu^i e^{-\tilde{q} \nu} P_k(\nu) d\nu \quad \text{(D7)}
\]

\[
H_k(\mu, \nu) = \int_{1}^{\infty} \int_{1}^{\infty} P_k (\mu(\nu)) Q_k (\mu(\nu)) \mu_1^i \mu_2^j e^{-w(\mu_1 + \mu_2)} d\mu_1 d\mu_2 \quad \text{(D8)}
\]

Here the \( P_k \) and \( Q_k \) are Legendre functions of the first and second kinds, respectively, \( \mu(\nu) \) is the smaller of \( \mu_1 \) and \( \mu_2 \), and \( \mu(\nu) \) is the larger of \( \mu_1 \) and \( \mu_2 \). The \( G_k \)
integrals were evaluated by Rosen (ref. 38). The \( H_k \) integrals were evaluated by James and Coolidge (ref. 39) and Rosen (ref. 40).

Two-center hybrid integral. - Here, three of the subscripts are identical. This case may be integrated by methods in reference 3. If the first three subscripts are identical,

\[
G_{mnop} = \frac{1}{2} R_{mp}^2 \sqrt{\zeta_m^3 \zeta_p^3} \left\{ \frac{e^{-w_{mp}}}{w_{mp}} \left[ \frac{e^{-\tilde{q}} - e^{-\frac{1 + w_{mp}}{\tilde{q}}} + e^{-\tilde{q}} \left( \frac{1}{\tilde{q}} + \frac{1}{\tilde{q}^2} \right) + e^{-\tilde{q}} \left( \frac{1}{\tilde{q}} - \frac{1}{\tilde{q}^2} \right)}{\tilde{q}} \right] \right. \\
- \frac{e^{-\tilde{r}}}{\tilde{r}} \left[ \frac{e^{-\tilde{q}} - e^{-\frac{1 + \tilde{r} + \tilde{r}^2}{\tilde{q}}} + e^{-\tilde{q}} \left( \frac{1}{\tilde{q}} + \frac{1}{\tilde{q}^2} \right) + e^{-\tilde{q}} \left( \frac{1}{\tilde{q}} - \frac{1}{\tilde{q}^2} \right)}{\tilde{q}} \right] \\
- \frac{1}{2} \zeta_m R_{mp} \frac{e^{-\tilde{r}}}{\tilde{r}} \left[ \frac{e^{-\tilde{q}} - e^{-\frac{2 + 2\tilde{r} + \tilde{r}^2}{\tilde{q}}} + e^{-\tilde{q}} \left( \frac{1}{\tilde{q}} + \frac{2}{\tilde{q}^2} + \frac{2}{\tilde{q}^3} \right) + e^{-\tilde{q}} \left( \frac{1}{\tilde{q}} + \frac{2}{\tilde{q}^2} - \frac{2}{\tilde{q}^3} \right)}{\tilde{q}} \right] \left\} \quad \text{(D9)}
\]

where \( w_{mp} \) is given by equation (3) and

\[
\tilde{q} = \frac{\zeta_m - \zeta_p}{2} R_{mp} \quad \text{(D10)}
\]
Three-center Coulomb integral. - Here, \( m = n \neq o \neq p \) or \( m \neq n \neq o = p \) in \( G_{mnop} \). The method of evaluation was the same as for equal orbital exponents.

Three-center exchange and four-center integrals. - The methods of evaluating these integrals were the same as for equal orbital exponents with the following exception. The methods of Magnasco and Dellepiane (ref. 37) require evaluation of

\[
U_{mn}(2) = \int \chi_m(1)\chi_n(1) \frac{1}{r_{12}} \, dv_1
\]

as a step in their procedure, where \( U_{mn}(2) \) is the electrostatic potential at electron 2 due to the bicentric charge distribution \( \chi_m(1)\chi_n(1) \). They give an expression for \( U_{mn}(2) \) if \( \zeta_m = \zeta_n \), but not if \( \zeta_m \neq \zeta_n \). A solution for \( \zeta_m \neq \zeta_n \) will now be given.

An infinite series for \( U_{mn}(2) \) with \( \zeta_m \neq \zeta_n \) may be derived by using spheroidal coordinates for both electrons 1 and 2. These coordinates were the same as in appendix C except that \( M \) replaced \( P \) and \( N \) replaced \( J \). The quantity \( 1/r_{12} \) was expressed as a Neumann expansion. The result is

\[
U_{mn}(2) = \frac{1}{2} R_{mn}^2 \sqrt{\zeta_m^3 \zeta_n^3} \sum_{k=0}^{\infty} (2k + 1) P_k(\nu_2) \left[ G_k^0(\tilde{\zeta}) K_k^2(w_{mn}, \mu_2) \right.
\]

\[
- G_k^2(\tilde{\zeta}) K_k^0(w_{mn}, \mu_2) + P_k(\mu_2) \left[ G_k^0(\tilde{\zeta}) L_k^2(w_{mn}, \mu_2) - G_k^2(\tilde{\zeta}) L_k^0(w_{mn}, \mu_2) \right] \right]
\]

where

\[
K_k^i(w, \mu_2) = \int_1^{\mu_2} e^{-w \mu_1} P_k(\mu_1) \mu_1^i d\mu_1
\]

and

\[
\bar{r} = \frac{3\zeta_m + \zeta_p}{2} R_{mp}
\]

\[
\bar{q} = \frac{3\zeta_m - \zeta_p}{2} R_{mp}
\]
The integrals \( K_k^j \) may easily be evaluated explicitly. To do this, the Legendre function of the first kind may be expressed as

\[
P_k(\mu) = \sum_{j=0}^{k} p_k^j \mu^j
\]  

(D17)

where the \( p_k^j \) are constants. Substituting equation (D17) into (D15) gives

\[
K_k^j(w, \mu) = \sum_{j=0}^{k} p_k^j A(w, \mu, j + i)
\]  

(D18)

where

\[
A(w, \mu, 0) = \frac{1}{w} (-e^{-w} \mu + e^{-w})
\]  

(D19)

and the recursion relation is

\[
A(w, \mu, h) = \frac{1}{w} \left[ -\mu^h e^{-w} \mu + e^{-w} + hA(w, \mu, h - 1) \right]
\]  

(D20)

The integrals \( I_k^j \) are not so easy to evaluate to six significant digits. The Legendre function of the second kind may be expanded in a hypergeometric series

\[
Q_k(\mu) = \frac{k!}{\mu^{k+1}(2k+1)!} \left[ 1 + \frac{\left( \frac{k+1}{2} \right) \left( \frac{k+3}{2} \right)}{\mu^2} + \frac{\left( \frac{k+1}{2} \right) \left( \frac{k+3}{2} \right)}{\mu^4} \right. \]

\[
+ \frac{\left( \frac{k+1}{2} \right) \left( \frac{k+3}{2} \right)}{\mu^6} \cdot \frac{\left( \frac{k+5}{2} \right)}{6 \mu^8} + \ldots \right]
\]  

(D21)
which is abbreviated as

\[ Q_k(\mu) = \frac{k!}{\mu^{k+1}} \sum_{j=0}^{\infty} \binom{k + 1}{2j} \binom{k + 1}{2j} j! \mu^{2j} \]  \hspace{1cm} (D22)

Substituting equation (D22) into (D16) gives

\[ L_k^i(w, \mu) = \frac{k!}{(2k + 1)!} \sum_{j=0}^{\infty} \frac{\binom{k + 1}{2j} \binom{k + 1}{2j}}{j!} \binom{k + 3}{2j} j! T(w, \mu, i-2j-k-1) \]  \hspace{1cm} (D23)

where

\[ T(w, \mu, h) = \int_{\mu}^{\infty} e^{-wu} u^h du \]  \hspace{1cm} (D24)

and \( h \) can be a positive or negative integer or zero. Equation (D24) was integrated analytically to get \( T \) in terms of exponentials and integral logarithms. If \( \mu \geq 1.2 \), equation (D23) gives six significant digits without too many terms. For \( \mu \leq 1.2 \), it is better to use

\[ L_k^i(w, \mu) = \int_{\mu}^{1.2} e^{-wu} Q_k(u)u^i du + L_k^i(w, 1.2) \]  \hspace{1cm} (D25)

where the first and second terms on the right were evaluated by Gaussian quadrature and equation (D23), respectively.

**General.** All \( G_{mnop} \) were evaluated to within \( \pm 4 \) in the sixth significant digit.
REFERENCES


26. Spindler, R. J., Jr.: Franck-Condon Factors for Band Systems of Molecular Hydrogen I. The \( \left( B \Sigma_u^+ - X \Sigma_g^+ \right) \), \( \left( \Pi_u^I - B \Sigma_u^+ \right) \) and \( \left( d \Pi_u^3 - a \Sigma_u^+ \right) \) Systems. J. Quant. Spectrosc. Radiat. Transfer, vol. 9, no. 5, May 1969, pp. 597-626.


"The aeronautical and space activities of the United States shall be conducted so as to contribute . . . to the expansion of human knowledge of phenomena in the atmosphere and space. The Administration shall provide for the widest practicable and appropriate dissemination of information concerning its activities and the results thereof."

—National Aeronautics and Space Act of 1958

NASA SCIENTIFIC AND TECHNICAL PUBLICATIONS

TECHNICAL REPORTS: Scientific and technical information considered important, complete, and a lasting contribution to existing knowledge.

TECHNICAL NOTES: Information less broad in scope but nevertheless of importance as a contribution to existing knowledge.

TECHNICAL MEMORANDUMS: Information receiving limited distribution because of preliminary data, security classification, or other reasons.

CONTRACTOR REPORTS: Scientific and technical information generated under a NASA contract or grant and considered an important contribution to existing knowledge.

TECHNICAL TRANSLATIONS: Information published in a foreign language considered to merit NASA distribution in English.

SPECIAL PUBLICATIONS: Information derived from or of value to NASA activities. Publications include conference proceedings, monographs, data compilations, handbooks, sourcebooks, and special bibliographies.

TECHNOLOGY UTILIZATION PUBLICATIONS: Information on technology used by NASA that may be of particular interest in commercial and other non-aerospace applications. Publications include Tech Briefs, Technology Utilization Reports and Notes, and Technology Surveys.

Details on the availability of these publications may be obtained from:

SCIENTIFIC AND TECHNICAL INFORMATION DIVISION
NATIONAL AERONAUTICS AND SPACE ADMINISTRATION
Washington, D.C. 20546