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INTERACTION OF NaCl(g) AND HCl(g) WITH CONDENSED Na_2SO_4

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ABSTRACT

The interaction of Na_2SO_4(1) with NaCl(g), HCl(g) and H_2O(g) was studied in atmospheric pressure flowing air and oxygen at Na_2SO_4(1) temperatures of 900 and 1000°C. Thermomicrogravimetric and high pressure mass spectroscopic sampling techniques were used. Experimental results establish that previously reported enhanced rates of weight loss of Na_2SO_4(1) in the presence of NaCl(g) are due to the reaction

\[ \text{Na}_2\text{SO}_4(\text{c}) + 2\text{HCl}(g) = 2\text{NaCl}(g) + \text{SO}_2(g) + \text{H}_2\text{O}(g) + 1/2\text{O}_2(g) \]

being driven to the right in flowing gas systems. The HCl(g) is the product of hydrolysis of NaCl caused by small but significant amounts of H_2O(g) present in the system. Thermochemical calculations are used to show that even with sub-ppm levels of H_2O(g) present, significant quantities of HCl(g) are produced.

INTRODUCTION

The role that condensed or gas phase sodium chloride plays in the hot corrosion process is not well understood and is the subject of a number of current investigations. When studying the effect of NaCl(c) on Na_2SO_4(1)-induced attack of metals, investigators use test methods in which mixtures of the salts are applied as deposits on samples or in which samples are immersed in crucibles containing mixtures of the salts. Maintaining the desired composition of the salt mixture is a problem in long duration, high temperature tests because the NaCl vaporizes preferentially. This is due to the fact that the vapor pressure of sodium chloride is orders of magnitude greater than that of sodium sulfate.

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Pettit and coworkers (refs. 1, 2) attempted to avoid the preferential loss of NaCl from NaCl(c)-Na2SO4(c) mixtures by adding partial pressures of NaCl(g) to the flowing air used in their tests. They reported that instead of suppressing the loss of NaCl from the mixture, the addition of NaCl(g) resulted in an increased rate of weight loss of the entire deposited mixture. This unexpected behavior prompted them to perform additional experiments with inert metal samples to eliminate any possible corrosion effects. Deposits of Na2SO4(c) were applied to platinum samples which were then heated to 1650°F in flowing air containing 300 ppm by weight of NaCl(g). Under these conditions, the samples lost weight at a linear rate of 0.4 mg cm^-2 hr^{-1}. Furthermore, the same rate of weight loss was observed when the inlet air was bubbled through water. In the absence of the NaCl(g) addition to the air flow, the rate of weight loss was reported to be 0.004 mg cm^-2 hr^{-1}. Thus the enhancement in rate of weight loss of Na2SO4(l), produced by the addition of NaCl(g) + H2O(g) to the air flow, was a factor of 100.

Jones et al. (ref. 3) carried out similar experiments with Na2SO4(c) deposits on platinum samples. With sample temperature held at 1600°F and 300 ppm NaCl(g) added to flowing air, they reported a weight loss rate of 0.03 to 0.05 mg cm^-2 hr^{-1} when the inlet air was passed through Drierite (anhydrous CaSO4) and 0.1 mg cm^-2 hr^{-1} when the air was bubbled through water.

Smeggil observed thermogravimetric results in his hot corrosion studies (ref. 4) which suggest that NaCl(g) promoted the loss of Na2SO4(l) from his samples. He performed additional experiments with Na2SO4(c) deposits on platinum substrates which he heated in flowing air with partial pressures of NaCl(g). Preliminary results indicated anomalously large weight loss rates (ref. 5).

The observations of Pettit et al., Jones et al. and Smeggil appear to be inconsistent with what one would predict on the basis of equilibrium thermodynamics. This fact and our continuing interest in hot corrosion phenomena prompted us to initiate the work reported herein. Our high pressure mass spectrometric sampling facility (ref. 6) is a powerful tool for studying systems at atmospheric pressure. Na2SO4(l)-NaCl(g) interactions were studied at a total pressure of one atmosphere of air or oxygen for various temperatures of Na2SO4(l) and for various partial pressures of NaCl(g) and H2O(g). Mass spectrometric sampling was used to identify and monitor gas phase species. In addition, continuous recording thermomicrogravimetric
measurements were used to determine condensed phase weight change rates. Experimental measurements were supplemented with thermodynamic calculations. Although this work is still in progress, the results presented here elucidate the interaction of NaCl(g) with Na\textsubscript{2}SO\textsubscript{4}(l).

EXPERIMENTAL

Materials. — For all of the experimental results reported here, Na\textsubscript{2}SO\textsubscript{4}(c) coatings were applied to Pt 20% Rh metal coupon substrates. The coupons were approximately 2.2 x 1.3 x 0.038 cm in size and had a 0.32 cm diameter hole near one end to facilitate hanging. Coatings of Na\textsubscript{2}SO\textsubscript{4}(c) were applied by spraying a saturated aqueous sodium sulfate solution onto the faces of the coupon as it was heated on a hot plate to about 150°C. The quantity of Na\textsubscript{2}SO\textsubscript{4}(c) applied was determined by weighing and ranged from 2 to 4 mg cm\textsuperscript{-2}. Fisher certified grade sodium sulfate was used.

Initially, gold coupons were tested as substrates for Na\textsubscript{2}SO\textsubscript{4}(c) but recorded weight change curves for various conditions were found to be not reproducible. Visual examination, after interrupting the experiment at various stages, revealed that the Na\textsubscript{2}SO\textsubscript{4}(c) coating after melting was not uniform over the sample. Distinct droplets of Na\textsubscript{2}SO\textsubscript{4}(c) were apparent and the size and distribution of the droplets were different for each test. Examination of a gold sample by scanning electron microscopy with energy dispersive X-ray analysis revealed that the Na\textsubscript{2}SO\textsubscript{4}(c) did not cover the entire surface area with even a thin film. Apparently Na\textsubscript{2}SO\textsubscript{4}(l) does not wet gold. Such a situation presents an unknown and nonreproducible surface area of Na\textsubscript{2}SO\textsubscript{4}(l) for vaporization and interaction with NaCl(g) during our tests. This then probably accounts for the lack of reproducibility observed in weight change curves for gold substrates.

Visual examination of Pt 20% Rh coupons coated with Na\textsubscript{2}SO\textsubscript{4}(c) was made at various stages of thermogravimetric tests. It was always noted that Na\textsubscript{2}SO\textsubscript{4}(l) tended to run down and form a thick section near the bottom of the coupon. Also, some Na\textsubscript{2}SO\textsubscript{4}(l) usually crept up the hangdown suspension hook. Thus, while the Na\textsubscript{2}SO\textsubscript{4}(l) appeared to wet the entire coupon surface area, we feel there is still a problem which precludes normalizing weight changes in terms of sample area.

Thermogravimetry. — Weight change experiments were carried out in the apparatus shown schematically in figure 1. A Cann type
RH microbalance, the output of which was continuously measured with a strip chart recorder, was used to follow the coupon weight change. The balance was operated with a sensitivity setting such that 5 micrograms of weight change produced one millimeter of deflection on the recorder chart. Because of noise produced in our flowing gas system, this was the highest practicable sensitivity that could be used.

The sample was hung from the balance with a 0.127 mm diameter hangdown wire with a 0.5 mm diameter wire hook at its lower end. The hangdown wire and hook were Pt 13% Rh alloy. The sample hung in a 2.3 cm inside diameter quartz tube attached to the balance enclosure by a greaseless ball and socket joint. To protect the balance from corrosive vapors, air was passed down through the balance enclosure at a rate of 365 cc/min. An equal flow of air was passed over the sample by flowing air into the inlet at the lower end of the quartz tube. Both gas flows exited from the system through the single gas exhaust shown in figure 1. Exhaust gas was bubbled through a 2.5 cm high column of vacuum pump oil. Air was supplied from the laboratory compressed air service line and each flow was passed through a 6 cm diameter by 20 cm high column of Drierite. When desired, the air flow over the coupon could, by appropriate valving, be bubbled through water before entering the system.

A portion of the quartz tube was located in a 2.86 cm inside diameter by 20 cm high movable furnace. The furnace temperature was controlled by a precision setpoint, proportional controller and silicon controlled rectifier power supply. Sample temperature was measured by a Pt-Pt 13% Rh thermocouple brought in through the bottom of the quartz tube. The thermocouple junction was located adjacent to the center of the sample. Ice-point compensated thermocouple voltage was read with a digital voltmeter.

Partial pressures of NaCl(g) were generated by heating pre-melted sodium chloride contained in a platinum crucible. The crucible was supported on a double bore alumina tube containing a thermocouple whose junction was in contact with the crucible bottom. The alumina tube-crucible assembly could be moved vertically through a slip-seal in the bottom end of the quartz tube. Crucible temperature was varied by vertical positioning of the crucible in the furnace and measured with the ice-point compensated thermocouple and a digital voltmeter. Sample and crucible temperatures were set and controlled to ±1°C.
The operating procedure was to hang the Na$_2$SO$_4$(c)-coated sample from the balance with the furnace cool and lowered below the sample. The salt crucible was lowered below the furnace. Gas flow was initiated and the balance was allowed to settle down to a steady state. The furnace was turned on and once it reached the set-point temperature, it was raised up to the desired position to heat the sample. The sample reached the operating temperature, either 900 or 1000°C, within five minutes after raising the furnace.

Prior to investigating Na$_2$SO$_4$(l)-NaCl(g) interactions, a number of "calibration" experiments were performed. The rate of weight gain due to condensation of NaCl on the hangdown wire was measured. This rate, $R$, in mg hr$^{-1}$ was measured, with an uncoated Pt 20% Rh coupon, as a function of coupon temperature and NaCl(g) concentration. The air flow was held constant at 365 cc min$^{-1}$ and a fixed position for the furnace was maintained. For coupon temperatures of 900, 950 and 1000°C, $R$ was found to be linear in time, independent of sample size and dependent on salt crucible temperature, $T_{NaCl}$, according to the Arrhenius relationship:

$$\log R = \frac{-A}{T_{NaCl}} + B$$

where $A$ and $B$ are constants whose value depends on the sample temperature.

The concentration of NaCl(g) added to the air flow was measured as a function of temperature of the salt crucible. Concentration, $C$ in ppm by weight, was determined by measuring the weight loss of salt in the crucible, after fixed increments of time at temperature, and calculating the weight of air that flowed in the time increment. For the flow rate of 365 cc/min, $C$ was found to vary with crucible temperature according to the relationship:

$$\log C = \frac{-a}{T_{NaCl}} + b$$

where $a$ and $b$ are constants whose value depends on the geometry of the apparatus.

High Pressure Mass Spectrometric Sampling.— A high pressure, free jet expansion, modulated molecular beam, mass spectrometric sampler and flow-tube system was used to identify directly gas phase species involved in Na$_2$SO$_4$(c)-NaCl(g) interactions under
various conditions. The apparatus is shown schematically in figure 2. Details of this apparatus have been presented previously (refs. 7, 8). The sample was suspended from a platinum wire in a 2.5 cm diameter quartz tube contained in a tube furnace. A smaller diameter, 1.5 cm, inner quartz tube extended up to within about a centimeter of the sample. Gas (oxygen or air) was fed independently to each tube and the total flow was 365 cc/min. The two-tube arrangement was used to facilitate adding water vapor to the outer tube only so as to protect the salt, contained in the Pt 20% Rh crucible, from hydrolysis. Sample and crucible temperatures were measured as in the thermogravimetric apparatus. All experiments were performed at atmospheric pressure.

The theory and description of the sampler have been presented elsewhere (ref. 6). Briefly, gas enters the sampler through a 0.25 mm diameter Pt 20% Rh sampling orifice located in the apex of a short 106 degree sampling cone. The gas is expanded as a free jet, converted to a molecular beam and passed, via differentially pumped vacuum stages, into a quadrupole mass spectrometer operated at pressures below 10^-8 torr. The molecular beam is modulated by a rotating segmented disc and synchronous detection is used to measure ion currents resulting from ions produced by electron bombardment of the neutral molecular beam.

When sampling systems containing condensable gas-phase species, it is desirable to have the sampling orifice as hot as practical to minimize clogging of the orifice. In our apparatus, the sampling cone protrudes into the flow tube and furnace as shown in figure 2. The orifice temperature attained in any experiment depends on the selected sample temperature. At sample temperatures below about 1000°C, with NaCl(g) concentrations above about 200 ppm, orifice clogging becomes a problem and limits the duration of experiments. Therefore, most experiments were performed at sample temperatures of 1000°C or higher and with NaCl(g) concentrations below about 150 ppm.

For the majority of the sampling experiments, oxygen was used as the flowing gas because it results in lower "metastable noise" (ref. 9) and produces "quieter" mass spectra than does air. However, even with the high purity tank oxygen used, interference is encountered in certain regions of the mass spectrum. Rare gas impurities contained in the oxygen are the cause of such interference. For example, the 40Ar+ signal precludes measurement of the NaOH+ ionic species.
RESULTS AND DISCUSSION

Thermogravimetric Experiments. — Figure 3 shows a typical recorded weight change curve for a Na$_2$SO$_4$(c)-coated Pt 20% Rh coupon at 900°C. Also shown in this figure is the linear weight gain curve (obtained for an uncoated coupon at 900°C when 300 ppm NaCl(g) was added to the air flow) that results from NaCl(c) pickup by the hangdown wire. Before NaCl(g) is added to the flow, the coated coupon is seen to lose weight linearly in time due to vaporization of Na$_2$SO$_4$(l). When a partial pressure of NaCl(g) (300 ppm by weight) is added to the "dry" air flow, a rapid weight gain is observed before the curve gradually levels off and transcends to a linear rate of weight loss. Finally, when H$_2$O(g) is added to the NaCl(g)-containing air flow, the rate of weight loss is seen to increase. The perturbation shown when the H$_2$O(g) was added is due to the disturbance of the balance produced by switching from dry to wet air.

Recorded weight changes for coated coupons must be corrected for (1) the concurrent weight gain due to NaCl condensation on the hangdown and (2) Na$_2$SO$_4$(l) vaporization weight loss. The first correction is obtained from "calibration" experiments and the second is obtained from the beginning of the curve obtained before the addition of NaCl(g). When such corrections are made, we see that the net weight loss rate resulting from the addition of NaCl(g) and NaCl(g) + H$_2$O(g) is appreciable.

The initial weight gain observed upon addition of NaCl(g) was presumed to be due to the absorption of NaCl by the Na$_2$SO$_4$(l). To verify this presumption, several experiments were interrupted during the initial weight gain by rapidly lowering the furnace and crucible and cooling the sample. The sample was removed from the apparatus and rinsed with water to remove all deposit. The resulting solution was analyzed for chloride. The calculated amount of NaCl, corresponding to the quantity of chloride, was found to be in excellent agreement with the respective recorded weight gain. These tests established that indeed the observed initial weight gain upon the addition of NaCl(g) is due to the absorption of NaCl by the Na$_2$SO$_4$(l). This point will be discussed in more detail later.

Numerous experiments were performed at sample temperatures of 900 and 1000°C with 300 ppm NaCl(g). In these experiments, the reproducibility of the Na$_2$SO$_4$ vaporization weight loss rate and initial weight gain upon addition of NaCl(g) were found to be
satisfactory. However, the weight loss rates measured after
the addition of NaCl(g) and H_2O(g) showed poor quantitative
reproducibility. At the present time, we cannot explain the
quantitative variation observed in our experimental results.
Nevertheless, we feel the qualitative features observed in
the thermogravimetric experiments are significant. Therefore,
the results of a typical set of experiments are presented in
Table I to serve as a qualitative summary of our observations.
These results show that the addition of NaCl(g) to air flowing
over Na_2SO_4(1) at 900 and 1000°C enhances the rate of weight
loss of the Na_2SO_4(1). This enhancement increases when H_2O(g)
is also added to the air flow.

Our results are in qualitative agreement with the results of
Jones et al. (ref. 3), but our findings regarding the addition
of H_2O(g) do not agree with the reported observations of Pettit
and co-workers (ref. 1). It is also interesting to note that
neither of these groups of investigators reported any observations
regarding the absorption of NaCl by Na_2SO_4(1) exposed to partial
pressure of NaCl(g).

The phase diagram for the NaCl-Na_2SO_4 system has been
measured and found to show a single eutectic with no soill solution
(ref. 10). Above 900°C, all compositions are completely liquidus.
Additionally, Flood et al. (ref. 11) have established that NaCl(1)
and Na_2SO_4(1) form a practically ideal solution in both
branches of the liquidus. Thus the quantity of NaCl(1)
absorbed by Na_2SO_4(1) in the presence of NaCl(g) can be calcu-
lated. Consider Na_2SO_4(1) as the solvent, designated by the
subscript S, and NaCl(1) as the solute, designated by the sub-
scripts. The weight of solute in the solvent is given by:

\[ W_s = \frac{X_s - M_s}{1 - X_s} \times \frac{W_C}{M_s} \]  

where W is weight, M is molecular weight and X is mole fraction.

Because sodium chloride in the gas phase is made up mainly
of monomer and dimer (for all practical purposes, higher order
polymers can be ignored), the mole fraction of solute can be
determined by considering the equilibria:

\[ \text{NaCl}(1) = \text{NaCl}(g) \]  

\[ 2\text{NaCl}(1) = (\text{NaCl})_2(g) \]
If we designate monomer and dimer by the subscripts m and d, respectively, the equilibrium constants for reactions (4) and (5) are:

\[ K_4 = \frac{P_m}{X_s^m} = P_m^o \]  
\[ K_5 = \frac{P_d}{X_s^d} = P_d^o \]

where \( P \) is pressure and the superscript \( o \) refers to the saturated equilibrium vapor pressure of pure NaCl(1). An effective pressure \( P_E \) can be defined as:

\[ P_E = P_m + 2P_d \]  

so that by substitution we obtain:

\[ P_E = X_s P_m^o + 2X_s^2 P_d^o \]

from which \( X_s \) can be determined by the quadratic formula giving:

\[ X_s = \left( \frac{(P_m^o)^2 + P_E}{16(P_d^o)^2} \right)^{\frac{1}{2}} - \frac{P_m^o}{4P_d^o} \]  

Thus we see that the mole fraction of solute is a function of the effective pressure of sodium chloride gas and the equilibrium saturated vapor pressures of the monomer and dimer. Because the equilibrium pressures are a function of temperature, the mole fraction of solute is also a function of the solute-solvent mixture temperature.

For our experiments, the effective pressure is related to the weight concentration, \( C \), of sodium chloride added to the air flow, i.e.:

\[ P_E = C \frac{M_{air}}{M_{NaCl}} \]  

\[ (11) \]
so that by substitution we obtain:

\[
\chi_5 = \left( \frac{(p_m^0)^2}{16(p_d^0)^2} + \frac{C_{\text{air}}}{2p_d^0 NaCl} \right)^{\frac{1}{2}} - \frac{p_m^0}{4p_d^0}
\]  

(12)

From equations (3) and (12), we see that the weight of NaCl(l) absorbed is a function of the weight of Na\textsubscript{2}SO\textsubscript{4}(l) present, the sample temperature and the concentration of NaCl(g) added to the air flow. In figure 4, we have plotted the mole fraction of NaCl(l) in Na\textsubscript{2}SO\textsubscript{4}(l) as a function of NaCl(g) concentration and mixture temperature. Values of the monomer and dimer equilibrium saturated vapor pressures were taken from the JANAF Thermochemical Tables (ref. 12).

For each of our experiments, the measured weight of NaCl(l) absorbed by the Na\textsubscript{2}SO\textsubscript{4}(l) was compared with the corresponding calculated value. As shown for the examples in Table I, the agreement was found to be excellent.

The absorption of NaCl(l) by Na\textsubscript{2}SO\textsubscript{4}(l) deposits exposed to NaCl(g) could have important ramifications for hot corrosion processes. When sodium chloride enters a turbine engine combustor, some will react with sulfur impurities in the fuel to form Na\textsubscript{2}SO\textsubscript{4}(g) (refs. 13, 14) which can, in turn, condense on hot gas path parts to form Na\textsubscript{2}SO\textsubscript{4}(l) (ref. 15). However, thermodynamic considerations predict (ref. 13) that NaCl(g) will also be present because all the sodium chloride is not converted to sulfate. Therefore, NaCl(l) could be absorbed by the Na\textsubscript{2}SO\textsubscript{4}(l) deposit. To assess how important this aspect might be, we have made some calculations for a burner rig test. The test itself is described elsewhere (ref. 15). For a burner rig flame doped with 8 ppm NaCl, 14.8 mg of Na\textsubscript{2}SO\textsubscript{4} were collected on a platinum target at 900°C. Equilibrium calculations for this flame condition (fuel/oxidant ratio = 0.0432, Jet A fuel, 8 ppm NaCl) give 5.62x10\textsuperscript{-7} and 6.14x10\textsuperscript{-11} atm, respectively, for the pressures of sodium chloride monomer and dimer. For the effective pressure of NaCl resulting from these pressures, we calculate the mole fraction of NaCl(l) absorbed as 2.8x10\textsuperscript{-4}. Thus for the 14.8 mg Na\textsubscript{2}SO\textsubscript{4}(l) deposit, the quantity of NaCl(l) absorbed is 1.7 µg. Actual analysis of the collected deposit from this test revealed that the chloride content was not
perceptible above the 40 μg chloride background level.

This calculation and comparison with a burner rig experiment indicate that, for the concentrations of NaCl(g) expected in turbine engines and the quantities of Na2SO4 deposits typically measured in practice, the amount of NaCl(1) absorbed is low.

**Mass Spectrometric Sampling Experiments.** - A typical set of experimental results obtained with Na2SO4(c) on a Pt 20% Rh sample are shown in figure 5. The measured ion intensities are relative to the one atmosphere oxygen signal and no corrections (cross section, Mach-number focusing and mass discrimination) were applied to measured values. In the first experiment with just oxygen flowing over the Na2SO4(c), the only pertinent species detected were O2(g), H2O(g) and Na2SO4(g). The H2O(g) is derived from the 600 ppm by volume of water vapor contained in the tank oxygen used. The Na2SO4(g) results from the vaporization reaction:

$$Na_2SO_4(c) = Na_2SO_4(g)$$

(13)

The saturated vapor pressure of Na2SO4(g) at 1000°C is 4.23x10^{-7} atm (ref. 13). No dissociation products of Na2SO4 were detected; this is as expected because for the reaction:

$$Na_2SO_4(c) = 2Na(g) + SO_2(g) + O_2(g)$$

at one atmosphere oxygen pressure and 1000°C the partial pressures of Na(g) and SO2(g) are calculated to be 1.3x10^{-8} and 6.7x10^{-9} atm respectively and these levels are below the detection limit of our sampler.

When about 100 ppm by weight of NaCl(g) was added to the oxygen flow, numerous pertinent vapor species were identified. In addition to the H2O(g) and various sodium chloride species, the mass spectrum was found to be relatively rich in Na+, SO2+ and HCl+ ions. Na2SO4+ and Na3SO4+ ions were also measured. The Na+ ion probably all derives from fragmentation of NaCl(g) in the electron bombardment (30 ev electron energy) ion source. The Na2SO4+ ion results from ionization of the Na2SO4(g) vaporization product. The precursor of the Na2SO4+ ion is presumed to be NaClNa2SO4(g) which loses a chlorine by fragmentation upon ionization. At the present time, we have not established whether the molecule NaClNa2SO4(g) is the product of a gas-solid interaction or gas-phase reaction.

For our considerations here the important ions observed are
HCl\(^+\) and SO\(_2\)\(^+\). Because some H\(_2\)O(g) is present in the flowing gas system, we attribute the HCl\(^+\) ion to the small amount of HCl(g) produced in the reaction:

\[
NaCl(g) + H_2O(g) = NaOH(g) + HCl(g)
\]  \hspace{1cm} (15)

even though interference prevented us from measuring NaOH(g). The SO\(_2\)\(^+\) ion is produced by ionization of SO\(_2\)(g) resulting from the reaction:

\[
Na_2SO_4(c, g) + 2HCl(g) = 2NaCl(g, c) + SO_2(g) + H_2O(g) + 1/2 O_2(g)
\]  \hspace{1cm} (16)

When additional H\(_2\)O(g) was added to the NaCl(g)-O\(_2\)(g)-Na\(_2\)SO\(_4\)(c) system, the hydrolysis product HCl(g) and reaction product SO\(_2\)(g) each increased markedly.

Experiments were performed where only 0.2% HCl(g) was added to the oxygen flow. In these experiments, NaCl(g) and SO\(_2\)(g) were detected as products. At temperatures below about 1100°C, orifice clogging was a problem. This was due to the relatively high concentration of NaCl(g) produced from reaction (16).

Results qualitatively similar to those discussed above were obtained when air was substituted for oxygen as the flowing gas. With air the amount of reaction of Na\(_2\)SO\(_4\)(c) was increased. This was ascertained by measuring the SO\(_2\)\(^+\) signals with air and oxygen under the various conditions. For example, with NaCl(g) and H\(_2\)O(g) added to the gas flow, the SO\(_2\)\(^+\) intensity with air was about two times greater than with oxygen. The increased reaction in the presence of air is due to the lower oxygen pressure associated with the air system.

The mass spectrometric results allow us to now qualitatively interpret the thermogravimetric experiments observations. Na\(_2\)SO\(_4\)(c) in flowing oxygen alone vaporizes and the Na\(_2\)SO\(_4\)(c)-coated sample loses weight at a low rate in proportion to the sample temperature and flow conditions. When NaCl(g) is added to the gas flow, some of it is hydrolyzed by water contained in the gas and the HCl(g) produced reacts with Na\(_2\)SO\(_4\)(c) according to reaction (16). Under this condition, the sample loses weight by both vaporization and reaction of Na\(_2\)SO\(_4\)(c) and the rate of weight loss is greater than before the NaCl(g) was added. When more H\(_2\)O(g) is added to the gas flow, more HCl(g) is produced and more reaction of Na\(_2\)SO\(_4\)(c) takes place. Thus the addition of more H\(_2\)O(g) results in a greater rate of weight loss of Na\(_2\)SO\(_4\)(c). Weight losses are observed in these experiments because the products of the Na\(_2\)SO\(_4\)(c) reaction are continually
being swept away from the sample by the flowing gas.

The qualitative description and reactions presented here to explain observed high temperature interactions of NaCl(g) with Na2SO4(c) under oxidizing conditions depend on the presence of H20(g). It is therefore of interest to calculate the dependence of such interactions on pressure of H2O(g). The equilibrium chemical composition of the Na-O-S-H-Cl system over liquid 0.954 (mole fraction) Na2SO4-0.046 NaCl at 1200K was calculated as a function of pressure of H2O(g). This Na2SO4-NaCl solution composition is that which is in equilibrium with 300 ppm by weight of NaCl(g) at this temperature. Calculations were performed with a modified version of the NASA Chemical Equilibrium Composition computer program (ref. 16) for an air atmosphere of 0.21 atm O2-0.79 atm N2. The results of these calculations are presented in figure 6 and show that even in air at equilibrium with Drierite some HCl(g) and NaOH(g) are present.

We feel that it is extremely difficult to remove low levels of water from flowing air or oxygen. While we had previously shown that the level of water vapor in ambient laboratory air (-10^-2 atm) is reduced by passage through Drierite (ref. 6), we have measured the level of water vapor in the tank oxygen used here and shown that it is unaffected by passage through a Drierite column. Thus we contend that without very special precautions, any study of the Na2SO4-NaCl system should include considerations of the interactions of water vapor.

Although the equilibrium constant for reaction (16) is much less than one for the conditions of our experiments, it is not surprising that the addition of HCl(g) leads to the removal of Na2SO4. This is due to LeChatelier's principle; continuous removal of the gaseous products of a reaction will cause the equilibrium to shift to the right. However, the same principle would imply that the addition of NaCl(g) should suppress the formation of gaseous products. This apparent paradox shows that thermodynamic considerations alone cannot explain the behavior observed here for the Na2SO4-NaCl system. Kinetically, for our experimental conditions, small amounts of HCl(g) are obviously more effective than large amounts of NaCl(g).

One additional mass spectrometric experiment was carried out to determine the effect of the addition of excess SO2 to the NaCl(l)-Na2SO4(l)-H2O(g)-O2(g)-SO2(g) system. Gaseous NaCl, 54 ppm by weight, was added to the oxygen flowing over a Na2SO4(l)-coated coupon. The coupon temperature was held at 1273K and a sufficient time was allowed to elapse to assure that NaCl(l) dissolved in
the Na₂SO₄(l). For these conditions, the mole fraction of NaCl(l) was calculated to be 3.4x10⁻³. Finally, the flow of NaCl(g) was stopped and SO₂(g) was added to the O₂(g) flow. The SO₂(g) was about 0.2% by volume. The mass spectral measurements revealed that both HCl(g) and Cl₂(g) were present. The measured partial pressure of HCl(g) was about two orders of magnitude greater than the partial pressure of Cl₂(g). McKee et al. (ref. 17) reported for a similar experiment only the detection of Cl₂(g).

To supplement our experiment, equilibrium calculations of the type used to produce figure 6 were made. The results of these calculations, given in figure 7, show that HCl(g) is indeed predicted to be a more abundant species than Cl₂(g) for the level of water vapor found in the tank oxygen we used in the experiments. Furthermore, the results of the calculations indicate that the species HCl, Cl and ClO should be more abundant than Cl₂ even for levels of water vapor below the theoretical value for H₂O in equilibrium with Drierite.

CONCLUDING REMARKS

The removal of Na₂SO₄(l) by NaCl(g), observed in our experiments and those of Pettit et al., Jones et al. and Smeggil, results from reaction (16) being driven to the right in a flowing gas system. The HCl(g) originates from the hydrolysis of NaCl by small amounts of water vapor present in the system. The only new gas phase species identified in Na₂SO₄(l)-NaCl(g) gas interactions is NaClNa₂SO₄(g) and experimental evidence suggests that it plays only a minor role in this interaction. Quantitative differences reported in measured thermogravimetric results are probably due to variability of H₂O(g) content and Na₂SO₄(l) surface coverage.

The results reported here have implications for hot corrosion processes and establish the importance of even trace amounts of water in the systems considered. Additional experiments and analysis are needed to further elucidate the role of NaCl(g) and NaCl(c).
REFERENCES


5. J. G. Smeggil, private communication.


12. JANAF Thermochemical Tables. Dow Chemical Co., Midland, MI.


**TABLE I**

THERMOGRAVIMETRIC RESULTS FOR PLATINUM SAMPLE
(AREA = 5.5 cm²) COATED WITH 13 mg OF Na₂SO₄ AND
HEATED IN FLOWING AIR* (FLOW VELOCITY - 1.6 cm/sec)

<table>
<thead>
<tr>
<th>System</th>
<th>Steady State Rate of Sample Weight Change, mg/hr</th>
<th>Weight of NaCl Absorbed Initially, mg</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>Temp, °C</td>
<td></td>
</tr>
<tr>
<td></td>
<td>900</td>
<td>1000</td>
</tr>
<tr>
<td>Na₂SO₄ + Dry Air</td>
<td>-0.01</td>
<td>-0.18</td>
</tr>
<tr>
<td>Na₂SO₄ + Dry Air + NaCl(g)**</td>
<td>-.08</td>
<td>-.32</td>
</tr>
<tr>
<td>Experimental</td>
<td></td>
<td></td>
</tr>
<tr>
<td>Calculated</td>
<td></td>
<td></td>
</tr>
<tr>
<td>Na₂SO₄ + Wet Air + NaCl(g)**</td>
<td>-.14</td>
<td>-.87</td>
</tr>
</tbody>
</table>

*Dry Air - Air Flowed Through Column of Drierite.
Wet Air - Air Bubbled Through Column of Water at 25°C.
**300 ppm (by Weight) of NaCl(g).

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Figure 1. - Schematic of thermogravimetric apparatus.

Figure 2. - Schematic of high pressure mass spectrometric sampling apparatus.
Figure 3. - Thermogravimetric recording of weight change of platinum sample (area = 5.5 cm²) coated with 13 mg Na₂SO₄ and heated at 900°C in flowing gas (flow velocity = 1.6 cm sec⁻¹).
Figure 4. - Variation of mole fraction of NaCl(g) in Na₂SO₄(g) with concentration of NaCl(g) in air at three mixture temperatures.

<table>
<thead>
<tr>
<th>SYSTEM</th>
<th>ION</th>
<th>RELATIVE INTENSITY**</th>
<th>MOLECULE</th>
</tr>
</thead>
<tbody>
<tr>
<td>Na₂SO₄(c) + O₂(g) + NaCl(g) 1000°C</td>
<td>H₂O⁺</td>
<td>10⁻⁸ 10⁻⁷ 10⁻⁶ 10⁻⁵ 10⁻⁴ 10⁻³ 10⁻²</td>
<td>H₂O Na₂SO₄</td>
</tr>
<tr>
<td></td>
<td>Na⁺</td>
<td></td>
<td>Na⁺</td>
</tr>
<tr>
<td></td>
<td>NaCl⁺</td>
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<td>NaCl⁺</td>
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<tr>
<td></td>
<td>Na₂Cl⁺</td>
<td></td>
<td>Na₂Cl⁺</td>
</tr>
<tr>
<td></td>
<td>SO₃⁺</td>
<td></td>
<td>SO₃⁺</td>
</tr>
<tr>
<td></td>
<td>Na⁺</td>
<td></td>
<td>Na⁺</td>
</tr>
<tr>
<td></td>
<td>NaCl⁺</td>
<td></td>
<td>NaCl⁺</td>
</tr>
<tr>
<td></td>
<td>NaSO₄⁺</td>
<td></td>
<td>NaSO₄⁺</td>
</tr>
<tr>
<td>Na₂SO₄(c) + O₂(g) + NaCl(g) + H₂(g) 1000°C</td>
<td>H₂O⁺</td>
<td></td>
<td>H₂O</td>
</tr>
<tr>
<td></td>
<td>NaCl⁺</td>
<td></td>
<td>NaCl⁺</td>
</tr>
<tr>
<td></td>
<td>SO₃⁺</td>
<td></td>
<td>SO₃⁺</td>
</tr>
<tr>
<td></td>
<td>HCl⁺</td>
<td></td>
<td>HCl⁺</td>
</tr>
<tr>
<td></td>
<td>NaOH⁺</td>
<td></td>
<td>NaOH⁺</td>
</tr>
</tbody>
</table>
| **TANK OXYGEN H₂O CONTENT INDEPENDENTLY ANALYZED AS 600 ppm BY VOL.
**MEASURED ION INTENSITIES (UNCORRECTED) RELATIVE TO 1 atm O₂(g) = 1.

Figure 5. - Mass spectral analysis of major vapor species emanating from a Na₂SO₄-coated platinum sample heated in flowing oxygen with and without impurity additions.
Figure 6. - Equilibrium composition of Na-O-S-H-Cl system over liquid 0.954 Na₂SO₄-0.046 NaCl at 1200 K.

Figure 7. - Equilibrium composition of Na-O-S-H-Cl system over liquid 0.9966 Na₂SO₄-0.0034 NaCl at 1273 K.