Metal phthalocyanine intermediates for the preparation of polymers

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Metal 4', 4'', 4'''-tetracarboxylic phthalocyanines (MPTC) are prepared by reaction of trimellitic anhydride, a salt or hydroxide of the desired metal (or the metal in powdered form), urea and a catalyst. A purer form of MPTC is prepared than heretofore. These tetracarboxylic acids are then polymerized by heat to sheet polymers which have superior heat and oxidation resistance. The metal is preferably a divalent metal having an atomic radius close to 1.35Å.

5 Claims, 3 Drawing Figures
FIG. 1
FIG. 2

\[ M = \text{Cu, Co, Ni, ...} \]
TGA & DTG CURVES OF THE SHEET POLYMERS OF COPPER PHTHALOCYANINE, COBALT PHTHALOCYANINE, NICKEL PHTHALOCYANINE.
METAL PHTHALOCYANINE INTERMEDIATES FOR THE PREPARATION OF POLYMERS

ORIGIN OF THE INVENTION

The invention described herein was made in the performance of work under a NASA contract and is subject to the provisions of Section 305 of the National Aeronautics and Space Act of 1958, Public Law 85-568 (72 Stat. 435; 42 U.S.C. 2457).

This is a division of application Ser. No. 415,880 filed Sept. 8, 1982, now U.S. Pat. No. 4,456,268, granted May 22, 1984.

FIELD OF THE INVENTION

This invention relates to metal phthalocyanine tetracarboxylic acids, to a novel method of synthesizing the same and to polymers of the same.

BACKGROUND OF THE INVENTION

In recent years there has been an increasing demand in industry and defense for polymeric substances which would either remain intact or continue to serve in a more or less degraded state under conditions where temperatures far above the ordinary are encountered. In many cases highly cross-linked organic polymers have been found promising. The phthalocyanine structure is one of the most thermally stable organic structures known. Attractive properties like resistance to chemical attack, electrical properties, catalytic activity and moderate cost of manufacture with good coloring properties have made phthalocyanines the object of intensive world-wide investigations. Many attempts to synthesize polymers based on phthalocyanines have failed to provide the expected thermal stability. Phthalocyanine polymers so far produced have shown less thermal stability than the phthalocyanine monomer itself because of the presence of impurities, weak chemical bonds and low degree of polymerization with structural inhomogeneity. Impurities have the considerable effect of decreasing thermal stability of the phthalocyanine monomers as well as the polymers produced from them. It is well known that the way in which the repeating units are linked is reflected in the properties of the polymers. If the repeating phthalocyanine mer units are linked in the way phenyl groups are linked in biphenyl, the conjugation extends throughout the macromolecule thereby increasing the extent of delocalization of the \( \pi \)-electrons. This is expected to increase the conductivity as well as the thermal stability of the phthalocyanine polymers.

OBJECTS OF THE INVENTION

Objects of the invention include methods of preparing metal phthalocyanine monomers in a pure state and the preparation from such monomers of thermally stable sheet polymers which have biphenyl type linkages between the mer units, high thermal and thermal oxidative stability and electrical conductivity, also good flame and fire resistance.

BRIEF DESCRIPTION OF THE INVENTION

A metal phthalocyanine tetracarboxylic acid (MPTC) is prepared in pure form by reaction of trimellitic anhydride with urea and a soluble metal salt in the presence of a suitable catalyst and in a suitable solvent. The metal is that which it is desired to introduce into the intended metal complex. The resulting MPTC, after separation and purification, is polymerized by heat in a vacuum or in an atmosphere of an inert gas such as nitrogen. The MPTC prepared by this method is purer than MPTC's prepared heretofore and the polymer derived from it is free of impurities which detract from its useful properties.
materials. The compound was re-precipitated with 1.0N hydrochloric acid and centrifuged to obtain the solid material. Dissolution and precipitation steps were repeated twice. Then the compound was washed until chloride free and finally washed with methanol. The blue precipitate was dried at 100°C in vacuum.

Compound C$_{36}$H$_{16}$N$_{8}$O$_{6}$Cu (CuPTC): Calcd: C, 57.5; H, 2.14; N, 14.9; Cu, 8.45; Eq. wt. 188.0. Found: C, 57.1; H, 2.3; N, 15.0; Cu, 8.5; Eq. wt. 187.6. IR absorption bands (cm$^{-1}$): 3500-2500 (broad), 1691 (broad), 1614m, 1578m, 1508m, 1329m, 1279m, 1246m (broad), 1188m, 1149m, 1089s, 1050w, 968w, 940w, 821w, 783m, 774w, 736s.

EXAMPLES 2 AND 3

The cobalt and nickel analogues of 1 were prepared by the same method using in Example 2 cobalt sulfate and in Example 3 nickel sulfate. The empirical formulae and analytical results obtained were as follows:

Example 2 (Cobalt)

Compound C$_{36}$H$_{16}$N$_{8}$O$_{6}$Co (CoPTC): Calcd: C, 57.8; H, 2.15; N, 15.0; Co, 7.88; Eq. wt. 186.9. Found: C, 57.5; H, 2.2; N, 15.2; Co, 7.90; Eq. wt. 186.7. IR absorption bands (cm$^{-1}$): 3500-2500 (broad), 1696 (broad), 1613m, 1586m, 1521m, 1330m, 1281m, 1246m (broad), 1189m, 1149m, 1090s, 1050w, 9730w, 944w, 848w, 782w, 773w, 742s.

Example 3 (Nickel)

Compound C$_{36}$H$_{16}$N$_{8}$O$_{6}$Ni (NiPTC): Calcd: C, 57.86; H, 2.15; Ni, 15.0; Ni, 7.86; Eq. wt. 186.8. Found: C, 57.91; H, 2.2; N, 15.05; Ni, 7.88; Eq. wt. 186.6. IR absorption bands (cm$^{-1}$): 3500-2500 (broad), 1699 (broad), 1615m, 1590m, 1530m, 1333m, 1274m, 1238m (broad), 1189m, 1150m, 1089s, 1050w, 976w, 944w, 848w, 782w, 779w, 738s.

Any mono-, di- or higher valency metal M may be used in place of divalent copper, cobalt and nickel, e.g., Cu(I), cobalt and nickel in various valence states, Fe(II), Fe(III), Zn, A1, lead, tin, palladium, germanium, vanium, platinum and molybdenum in various valence states; monovalent metals such as Li, Na and K, etc. Where the metal is monovalent two atoms will be present, one atom above and the other atom below the plane of the phthalocyanine molecule. With a trivalent metal such as aluminum, one of the valences may be satisfied by a ligand such as Cl, F, acetate, etc., e.g., the metal aluminum molybdate (which provides the benzenoid rings), urea and possibly also the ammonium chloride. The ammonium chloride may be used alone as the catalyst but the use of ammonium molybdate as well allows the reaction to proceed at a lower temperature and increases the yield.

Nitrobenzene is an advantageous solvent because, besides being unreactive and being a good solvent for the reactants, it has a high boiling point (210°C). Other aprotic solvents such as quinoline and tetralin may be used. Yields of 90% or more are obtainable.

The metal M may be removed from the molecule as by dissolving it in concentrated sulfuric acid to produce the hydrogen (protonated) species in which each of the two covalent bonds is connected to hydrogen. The protonated species may then be treated with an alcoholic solution of a metal salt, e.g., copper sulfate, to insert the metal, e.g., Cu(II). In this way one metal may be substituted for another. However, it is preferred to choose the desired metal initially and to insert it in the phthalocyanine molecule in the synthesis of the tetracarboxylic acid.

EXAMPLE 4

Polymerization of MPTC—Method 1

About 1-2 g of metal phthalocyanine tetracarboxylic acid was finely ground in a small vibrating ball mill and placed in a polymerization tube. The tube was connected to an apparatus provided with a tube to condense volatile products, stopcocks to connect gas collection tube and IR cell for gaseous analysis. The apparatus was connected to vacuum system. Reaction tube was carefully evacuated to 10$^{-2}$-10$^{-8}$ torr pressure and heated to 450°C. The gaseous and volatile products were condensed in separate traps using liquid nitrogen. The reaction was found to be completed after one hour of heating at 450°C in vacuum.

EXAMPLE 5

Polymerization of MPTC—Method 2

Finely ground metal phthalocyanine tetracarboxylic acid was taken in a reaction tube fitted with a glass enclosed iron-constantan thermocouple and inlet and outlet tubes. The polymerization tube was carefully purged with nitrogen by repeated evacuation and refilling. Then it was gradually heated to 450°C in a current of nitrogen and maintained temperature of 450-500°C in a current of nitrogen for one hour.

Methods 1 and 2 were applied to produce the divalent copper, cobalt and nickel compounds. Analytical results were the same for products of the two methods and were as follows:

(1) Compound C$_{36}$H$_{16}$N$_{8}$Cu: Calcd: C, 67.18; H, 2.11; Cu, 19.58; Eq. wt. 66.5. Found: C, 66.9; H, 2.21; Cu, 19.8; Eq. wt. 65.5. IR absorption bands (cm$^{-1}$): 1610w, 1503w, 1407w, 1328w, 1271m, 1164m, 1118w, 1089m, 1065m, 989w, 773w, 754w, 738s.

(2) Compound C$_{36}$H$_{16}$N$_{8}$Co: Calcd: C, 67.73; H, 2.13; Co, 19.75; Eq. wt. 67.6. Found: C, 67.6; H, 2.3; Co, 19.95; Eq. wt. 10.46. IR absorption bands (cm$^{-1}$): 1599w, 1502w,
These polymers are soluble in concentrated (98%) sulfuric acid, concentrated (36%) hydrochloric acid, chlorosulfuric acid and trimethyl sulfonic acid. They have high thermal resistance and have a higher electrical conductivity (as well as a higher thermal oxidative resistance) than the phthalocyanine monomers. Threshold temperatures at which major decomposition occurs are typically higher than 450° C. in air; they have a high char yield in nitrogen and in nitrogen atmosphere no catastrophic decomposition has been observed up to 1000° C. Char yields at 800° C. in nitrogen and electric conductivities are given in Table I.
4,522,755

TABLE 1

<table>
<thead>
<tr>
<th>POLYMER</th>
<th>( \text{PDT}_{\text{max}} ) °C</th>
<th>YIELD (22-300°C)</th>
<th>CONDUCTIVITY (10-cm)</th>
</tr>
</thead>
<tbody>
<tr>
<td>CuPc-Sheet</td>
<td>760</td>
<td>90.5%</td>
<td>3.2 \times 10^{-10}</td>
</tr>
<tr>
<td>Polymer</td>
<td>860</td>
<td>89.0%</td>
<td>1.8 \times 10^{-2}</td>
</tr>
<tr>
<td>CoPc-Sheet</td>
<td>890</td>
<td>93.0%</td>
<td>1.4 \times 10^{-3}</td>
</tr>
<tr>
<td>NiPc-Sheet</td>
<td>890</td>
<td>93.0%</td>
<td>1.4 \times 10^{-3}</td>
</tr>
</tbody>
</table>

In FIG. 3 of the drawing the results plotted are for Cu, Co and Ni sheet polymers heated in an atmosphere of nitrogen. The ordinate scale on the left represents percent decomposition and the ordinate scale on the right represents rate of weight loss. As will be seen the polymers did not undergo substantial decomposition below 750°C to 800°C. (See curves at top.) The rate of decomposition (lower set of curves) did not become substantial until about 750°C to 900°C.

It will be apparent that purer metal phthalocyanine tetracarboxylic acids are provided, that a new and useful method of preparing them has been provided and that new and useful polymers and methods of producing them have been provided.

It is claimed:

1. A method of preparing an isolated metal 4,4',4'',4'''-tetracarboxylic phthalocyanine which comprises
   a. heating at about 185°C for about four hours a reaction mixture comprising an anhydride consisting essentially of trimellitic anhydride, a source of metal selected from finely divided powders and salts of said metal, and urea in the presence of a catalyst comprising ammonium chloride and ammonium molybdate in a nitrobenzene reaction medium, and
   b. recovering the desired metal 4,4',4'',4'''-tetracarboxylic phthalocyanine from the reaction mixture by,
   b.1. washing the reaction mixture with methanol to give a washed product,
   b.2. boiling the washed product in aqueous mineral acid to give an acid-wash product,
   b.3. treating the acid wash product with inorganic hydroxide at elevated temperature until ammonia evolution ceases and a soluble product results,
   b.4. precipitating the desired metal 4,4',4'',4'''-tetracarboxylic phthalocyanine as a solid from the soluble product, and
   b.5. isolating and washing the solid metal 4,4',4'',4'''-tetracarboxylic phthalocyanine with methanol and drying.

2. The method of claim 1 wherein the metal is a divalent metal having an atomic radius of approximately 1.35 Å.

3. The method of claim 2 wherein the metal is copper and the source of metal is a copper salt.

4. The method of claim 2 wherein the metal is nickel and the source of metal is a nickel salt.

5. The method of claim 2 wherein the metal is cobalt and the source of metal is a cobalt salt.

* * * * *
UNITED STATES PATENT OFFICE
CERTIFICATE OF CORRECTION

Patent No. 4,522,755 Dated June 11, 1985

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It is certified that error appears in the above-identified patent and that said Letters Patent are hereby corrected as shown below:

First page, line [22], change "Sep. 15, 1983" to --July 15, 1983--.

First page, line [62], change "Pat. No. 4,456,268" to --Pat. No. 4,450,268--.

Column 1, line 12, delete "4,456,268" and insert --4,450,268--.

Signed and Sealed this First Day of October 1985

[SEAL]

Attest: DONALD J. QUIGG

Attesting Officer Commissioner of Patents and Trademarks—Designate