A comprehensive model to determine the effects of temperature and species fluctuations on reaction rates in turbulent reacting flows*

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INTRODUCTION

The research summarized herein was begun on February 1, 1980 and had a two-fold objective:

1. The development of a computationally-viable model describing the interaction between fluid-mechanical turbulence and finite-rate combustion reactions, principally in high-speed flows; and

2. The development of chemical kinetic mechanisms, complete and global, describing the finite rate reaction of fuels of interest to NASA with air. These fuels included principally hydrogen and silane, although a limited amount of work involved hydrocarbon fuels as well.

During the course of this research, a number of external publications were issued which describe the work accomplished in substantial detail. As a result, in this Final Report relatively little emphasis will be placed upon the accomplishments adequately described elsewhere. Instead, our emphasis here will be upon aspects of the work not adequately described in the open literature.
THE TURBULENCE - REACTION CHEMISTRY INTERACTION MODEL

This model is described in substantial detail in references 1 and 2 and is summarized in reference 3. Basically, the model is of the "assumed pdf" type in which turbulent mean reaction rate terms are expressed in the form

\[
\dot{\omega}_{A_{1},1} = \left[ \int_{-\infty}^{\infty} k_{f,i}(T)p(T) \,dT \right] \left[ \int_{-\infty}^{\infty} \int_{-\infty}^{\infty} c_{A_{1}} c_{A_{2}} \cdot p(c_{A_{1}}, c_{A_{2}}) \,dc_{A_{1}} \,dc_{A_{2}} \right]
\]

(1)

in which \(k_{f,i}\) is the forward reaction rate coefficient for a particular reaction in a kinetic mechanism, \(p(T)\) is the probability density function (pdf) of the temperature, \(C_{A_{j}}\) are species concentrations, and \(p(c_{A_{1}}, c_{A_{2}})\) is the joint pdf of these concentrations. Non-dimensionalizing the temperature and species concentrations and introducing the Arrhenius equation for \(k_{f,i}\) ultimately leads to (refs. 1 - 3)

\[
\dot{\omega}_{A_{1},1} = (Z_{t} k_{f,i}) Z_{r} \bar{c}_{A_{1}} \bar{c}_{A_{2}}
\]

(2)

in which \(Z_{t}\) is termed a temperature amplification ratio and \(Z_{r}\) is the species amplification ratio. Extensive work with a number of temperature pdf's revealed that values of \(Z_{t}\) are
relatively insensitive to its selection; hence, the computationally simple beta pdf was recommended for use. The recommended joint (two-variable pdf) required for the determination of $Z_r$ is the most likely pdf. Typically, $Z_t$ values appear as in Figs. 1 and 2, indicating that $Z_t$ always exceeds unity. Hence, the effect of the turbulent temperature fluctuations is to increase the reaction rate coefficients.

On the other hand, $Z_r$ may be greater than one (Fig. 3), an enhanced mixedness effect, or less than one (Fig. 4), an unmixedness effect, depending upon the sign of the species concentration gradient in the flow field. The assumed signs of those gradients depends upon the elementary reaction considered and upon the presumed location of the "flame front" location depicted in Fig. 5. Table I indicates the signs of the correlation coefficient $\rho^*$ discussed in ref. 2; $\rho^* > 0$ implies $Z_r > 1$, $\rho^* < 0$ implies $0 < Z_r < 1$. The results in ref. 2 assumed the "flame front" location to be at fuel-to-oxygen) equivalence ratio equal to one; ie, $\varnothing_{cr} = 1$. Subsequent work revealed that for the case examined in ref. 2, values in appreciably better agreement with the experimental values are obtained for $\varnothing_{cr} = 0.1$. These results are shown in Figs. 6 - 17.

A study was also undertaken to assess the effects of changes in the constants associated with the generation and dissipation terms in the transport equations discussed in references 4 and 5 and used in reference 2. In particular, the constants $C_{\varepsilon_1}$ and $C_{\varepsilon_2}$ in the equation for the dissipation of turbulence kinetic energy and $C_{g_1}$ and $C_{g_2}$ used in all species fluctuations transport
equations and in the transport equation for the temperature fluctuations. The values consistently used heretofore are:

\[
\begin{align*}
C_{\varepsilon_1} & = 1.43 \\
C_{\varepsilon_2} & = 1.92 \\
C_{g_1} & = 2.80 \\
C_{g_2} & = 2.00
\end{align*}
\]

with \( C_{\varepsilon_2} \) modified for the case of axisymmetric flow.

The study of \( C_{\varepsilon_1} \) and \( C_{\varepsilon_2} \) was prompted by the conflicting results of two sensitivity analyses each of which dealt with a different flow situation: one, performed on a \( k-\varepsilon \) turbulence model for incompressible round-jet flow showed large sensitivity to a one percent change in the empirical coefficients \( C_{\varepsilon_1} \) and \( C_{\varepsilon_2} \) in the \( \varepsilon \) transport equation [Ref. 6]; on the other hand, Pinckney's analysis on a \( k-\varepsilon \) turbulence model for "turbulent mixing of hydrogen injected from discrete holes in the surface of a rectangular duct" showed little sensitivity to a 10% change in \( C_{\varepsilon_1} \) and \( C_{\varepsilon_2} \) [Ref. 7].

For the experimental case in ref. 4, both \( C_{\varepsilon_1} \) and \( C_{\varepsilon_2} \) were increased by 1% as in Reference 6. The results* (Figs. 18 and 19) showed no sensitivity to the increase in the constants. That these results agree with Pinckney's findings further confirms the correctness of the standard values for \( C_{\varepsilon_1} \) and \( C_{\varepsilon_2} \) for turbulent, compressible flow.

* Note that the only parameters used to compare the results are \( H_2O \) and \( O_2 \) concentrations because pitot pressure and \( N_2 \) concentrations have been found to be insensitive to changes in the turbulence model [Ref. 6].
One of the transport equations used in the model is that for temperature fluctuations [Ref. 8]. This is treated analogously to the equations for the transport of fluctuations in fuel and oxidizer, as well as the dissipation rate of turbulent kinetic energy. The fluctuation equation for fuel and oxidizer are both dependent on the constants $C_{g1}$ and $C_{g2}$, the commonly recommended values of which are given above. However, no references have been located setting forth the equivalent constants (here called $C_1$ and $C_2$) in the temperature fluctuation equation; so the code uses $C_1 = C_{g1}$ and $C_2 = C_{g2}$.

An investigation was made of the assumed value of $C_2$. The model was tested for values of $C_2 = (0.9)C_{g1}$ and $C_2 = (0.1)C_{g2}$. The ten percent decrease in the value of $C_2$ did not cause any change in the results. However, decreasing the value by one order of magnitude caused significant changes (Fig. 20). Improvement was found from the centerline radially outward until the start of combustion. Here, the mass fraction of oxygen was found to be significantly lower and the mass fraction of water was found to be somewhat higher.

This work suggests that the value of $C_2$ in the equation for the transport of temperature fluctuations should perhaps not be the same as the value of $C_{g2}$ in the equations for fluctuations of fuel and oxidizer. It is likely that $C_1$, assigned the value of $C_{g1}$, should also be revised.

All the transport equations upon which this model is based require initial radial profiles. Initial values for the equation for the temperature fluctuations are required. There has been
no experimental or theoretical work done on this initial profile. In the code as employed in ref. 2, the initial value for the non-dimensional temperature fluctuations is $t' = 0.003$ for all radial locations at the initial axial location. This is also the value used for the initial profiles in the transport equations for the fluctuations of fuel and oxidizer.

To determine the effects of changing the initial profile, two attempts were made to improve this profile. First, the initial value was raised by one order of magnitude to 0.03 (Fig. 21). This resulted in some improvement near the centerline and slight improvement in the combustion zone.

Evans, et. al. (ref. 4) reported that the initial profiles used in the transport equations for turbulence kinetic energy and its dissipation had a strong influence on their results. Their report gave two possible initial $k$-$\varepsilon$ profiles: one by a theoretical method and the other deduced from experimental measurements and shown in Fig. 22.

The YCHARML code generates $k$-$\varepsilon$ profiles according to the method discussed in ref. 4. These profiles had been used in all previous work, so in this study the "experimental" profiles (next page) were tried using a $\phi_{cr}$ of 1.0 (Fig. 5) and 0.1 (Fig. 6). As before (ref. 6), it was found that the critical equivalence ratio of 0.1 gave the better results (Fig. 7). However, the experimental initial $k$-$\varepsilon$ profiles did not materially improve the results. From the centerline radially outward to the flame zone, the experimental profiles gave slightly better results, but outside the flame zone, the results were significantly worse.
CHEMICAL KINETICS STUDIES

Two principal results of these studies were the development of a global \( \text{H}_2 \) - air combustion model, described in reference 9, and a complete silane (\( \text{SiH}_4 \))-\( \text{H}_2 \)-air mechanism, described in reference 10. The proposed hydrogen-air global combustion model consists of the two "elementary" reactions

\[
\begin{align*}
\text{H}_2 + \text{O}_2 & \rightarrow \text{OH} + \text{OH} \quad \text{forward rate coefficient} \quad k_f^1 \quad \text{with} \quad N_1 = 10 \\
2 \text{OH} + \text{H}_2 & \rightarrow 2 \text{H}_2\text{O} \quad \text{forward rate coefficient} \quad k_f^2 \quad \text{with} \quad N_2 = -13
\end{align*}
\]

with the forward rate coefficients given in the form

\[
K_{f_i} = A_i(\phi)T^{N_i} \exp\left(-\frac{E_i}{RT}\right)
\]

(3)

The recommended values (ref. 9) at one atmosphere pressure are

\[
\begin{align*}
A_1(\phi) &= (8.917\phi + 31.4 \frac{33}{\phi} - 28.950) \times 10^{47}, \text{ cm}^3/\text{mol}^5 \\
E_1 &= 4865 \text{ cal/mol} \\
N_1 &= 10 \\
A_2(\phi) &= (2.0 + 1.333/\phi - 0.833\phi) \times 10^{64}, \text{ cm}^3/\text{mol}^5 \\
E_2 &= 42,500 \text{ cal/mol} \\
N_2 &= -13
\end{align*}
\]
To account for the effects of pressures between 0.5 atm and 1.0 atm, $A_2$ must be revised as follows:

For $\phi > 1$

$$A_2 = 8.80 + \frac{5.85}{\phi} - 3.67\phi - 4.80 \ p - 3.20 \ p/\phi + 2.00 \ p \ \phi$$

For $\phi < 1$

$$A_2 = 0.67 \ \phi^{-1.5} \ \ p^{-0.7} \ (8.80 + \frac{5.85}{\phi} - 4.80 \ p - 3.20 \ p/\phi + 2.00 \ p \ \phi)$$

The silane-hydrogen-air ignition and combustion studies are detailed in ref. 10.

Current interest in the use of hydrocarbon fuels in SCRAMJET engines requires that methods be sought for reducing the well-documented lengthy ignition delays and reaction times of hydrocarbon fuels. Two possible approaches to reducing these times are: (1) injecting (relatively) small quantities of hydrogen along with the hydrocarbon fuel with the hope that ignition delay times will be more like those for $H_2$ than for the $C_xH_y$; and (2) regeneratively heating the hydrocarbon prior to its injection into the combustion chamber so as to pyrolyze (thermally crack) it, with the expectation that substantial quantities of hydrogen will be among the pyrolysis products. Studies are underway to investigate both possibilities.

Calculated ignition delay times up to 33% (vol.) hydrogen in the fuel are shown in Fig. 26. Calculations have been carried out selecting propane ($C_3H_8$) as the representative hydrocarbon fuel.
As can be seen in Fig. 26, at 1.0 atm, $\phi = 1.0$, the reduction in ignition delay time compared with that for pure C$_3$H$_8$ is slight. For example, at $T_0 = 1000$K, $t_{ID} \approx 0.05$ sec for C$_3$H$_8$ and about 0.025 sec for the 33% H$_2$/67% C$_3$H$_8$ mixture. On the other hand, $t_{ID} = 10^{-4}$ sec for pure H$_2$ at $T_0 = 1000$K. It is recognized, however, that 33% H$_2$ in the fuel mixture represents very little hydrogen by mass. As an example, the mass equivalent of the 20% SiH$_4$/80% mixture as applied here would result in a fuel mixture of 98.88% H$_2$/1.12% C$_3$H$_8$ by volume. A similarly wide range of H$_2$/C$_x$H$_y$ ratios was utilized in the experiments of Cookson (ref. 11), shown schematically in Fig. 27. His tests, the results of which are in Fig. 28, would span the range from about 3% H$_2$ to nearly 96% H$_2$ (by volume) had his "main" fuel been C$_3$H$_8$ rather than kerosene. It is interesting to note that using mode A (see Fig. 27), ignition could be achieved at reasonable kerosene injection pressure ratios using relatively small quantities of H$_2$ (about 3% to 33% by volume), whereas mode B required much more substantial quantities of H$_2$ (33% - 96% by volume). Hence, calculations were carried out employing hydrogen concentrations exceeding 33% by volume. The effect upon $t_{ig}$ is shown in Fig. 29 ($\phi = 1$, $p = 1$ atm). It is clear that appreciable reductions of $t_{ig}$ do not occur for values of H$_2$ concentration less than 40%.

In order to arrive at a recommended minimum fuel hydrogen concentration for SCRAMJET engine applications, a cross-plot was made of the data in Fig. 29 and is shown in Fig. 30 at $T_0 = 1000$ K and 1667 K. It is clear, looking at the 1000K curve, that very large
reductions in $t_{ig}$ occur for fuel hydrogen concentrations exceeding 80\% by volume (15.4\% by weight). Hence, from the ignition viewpoint, fuel hydrogen concentrations exceeding that value are recommended.

Similar $t_{ig}$ behavior occurs for values of equivalence ratio other than stoichiometric, as is shown in Figs. 31 and 32. This tends to reinforce the above recommendation of $\% \text{H}_2 > 80\%$ (vol.).

From the viewpoint of reaction time, $t_R$, no clear effect of fuel hydrogen concentration was observed for the conditions calculated. The range of $t_R$ values observed for hydrogen concentrations ranging from 0\% to 96\% is shown in Fig. 33.

The flame stabilizing characteristics of various C$_3$H$_8$/H$_2$ fuels are shown in the blowout limit correlations in Fig. 34. Here again, as may be seen in the cross-plot at $\phi = 1$ in Fig. 35, values of fuel hydrogen concentration exceeding 80\% may be said to result in an "appreciable" shift of the stable flame envelope.

On the basis of these studies, fuel hydrogen concentrations exceeding 80\% by volume (15.4\% by weight) are recommended for SCRAMJET engines employing hydrogen - gaseous hydrocarbon fuel mixtures, provided that the fuel is propane or a higher molecular weight hydrocarbon. No a priori conclusions can be drawn concerning such lower molecular weight fuels as methane, ethane, ethene and acetylene.

In addition to the above studies, an extensive investigation was undertaken dealing with the ranges of applicability of chemical kinetic models of hydrogen-air combustion. A description of these studies is in Appendix A herein.


REFERENCES


<table>
<thead>
<tr>
<th>Reaction</th>
<th>Sign of the Correlation Coefficient</th>
<th>Forward</th>
<th>Reverse</th>
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<tbody>
<tr>
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<td>$\phi &gt; \phi_{cr}$</td>
<td>$\phi &lt; \phi_{cr}$</td>
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<td>1. $\text{H}_2 + \text{M} = \text{H} + \text{H} + \text{M}$</td>
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<td>3. $\text{H}_2\text{O} + \text{M} = \text{OH} + \text{H} + \text{M}$</td>
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<td>4. $\text{OH} + \text{M} = \text{O} + \text{H} + \text{M}$</td>
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<td>5. $\text{H}_2\text{O} + \text{O} = \text{OH} + \text{OH}$</td>
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<td>6. $\text{H}_2\text{O} + \text{H} = \text{OH} + \text{H}_2$</td>
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<td>7. $\text{O}_2 + \text{H} = \text{OH} + \text{O}$</td>
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<tr>
<td>8. $\text{H}_2 + \text{O} = \text{OH} + \text{H}$</td>
<td>-</td>
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</tbody>
</table>
Figure 1.- Comparison of results for beta pdf and most-likely pdf for case 1.
Figure 2

\( \bar{t} = 0.2 \) and \( 0.4 \).

\( \bar{t} = 0.2 \) and \( 0.4 \).
$\frac{r_{A_1}^2}{r_{A_2}^2} = 0.09$

$\bar{r}_{A_1} = 0.4; \bar{r}_{A_2} = 0.6.$

Figure 3
The figure shows a graph with the $Z_r$ on the y-axis and $\overline{r_{A_1}^2}$ on the x-axis. The graph includes several curves labeled with values $r_{A_1} = 0.2$ and $r_{A_2} = 0.9$. The title of the figure is 'Figure 4.'
Figure 5. Schematic of the Axisymmetric Flame
Figure 6. Radial species-concentration profiles for oxygen.
Figure 7 Radial species-concentration profiles for oxygen.
Figure 8  Radial species-concentration profiles for oxygen.
Figure 3. Radial species-concentration profiles for water.
Figure 11. Radial species-concentration profiles for water.
Figure 12. Radial pitot-pressure profiles.
Figure 13. Radial pitot-pressure profiles.
x/dj = 13.5  \hspace{1cm} \phi_i = 0.1

\begin{figure}[h]
\centering
\includegraphics[width=\textwidth]{figure14_radial_pitot_pressure_profiles.png}
\caption{Radial pitot-pressure profiles.}
\end{figure}
Figure 15. Radial pitot-pressure profiles.
Figure 16. Radial pitot-pressure profiles.
Figure 17  Radial pitot-pressure profiles.
Figure 18 Species concentration profiles for a one percent increase in $C_{e1}$. 

$8.26 \quad C_{e1} = 1.4443$

**Experimental**
- H$_2$O
- O$_2$

**Theoretical**
- $C_{e1} = 1.43$
- H$_2$O - $C_{e1} = 1.4443$
- O$_2$ - $C_{e1} = 1.4443$
15.4  $C_{e1} = 1.4443$

**Theoretical**
- $C_{e1} = 1.43$
- $H_2O - C_{e1} = 1.4443$
- $O_2 - C_{e1} = 1.4443$

**Figure 3. Continued.**
$21.6 \quad C_{e1} = 1.4443$

**Experimental**

- $\odot$ H$_2$O
- $\square$ O$_2$

**Theoretical**

- $\longrightarrow$ $C_{e1} = 1.43$
- $\times$ H$_2$O - $C_{e1} = 1.4443$
- $\rightarrow$ O$_2$ - $C_{e1} = 1.4443$

*Figure 1B. Concluded.*
8.26 \[ C'_{\varepsilon_2} = 1.01 \times C_{\varepsilon_2} \]

**Experimental**
- \( H_2O \)
- \( O_2 \)

**Theoretical**
- \( C_{\varepsilon_2} \)
- \( H_2O - C_{\varepsilon_2}' = (1.01)C_{\varepsilon_2} \)
- \( O_2 - C_{\varepsilon_2}' = (1.01)C_{\varepsilon_2} \)

**Figure 19** Species concentration profiles for a one percent increase in \( C_{\varepsilon_2} \).
15.4 \[ C'_{\text{CO}_2} = 1.01 \times C_{\text{CO}_2} \]

**Experimental**
- \( \bullet \) \( \text{H}_2\text{O} \)
- \( \bullet \) \( \text{O}_2 \)

**Theoretical**
- \( \text{C}_{\text{CO}_2} \)
- \( \times \) \( \text{H}_2\text{O} - C'_{\text{CO}_2} = (1.01)C_{\text{CO}_2} \)
- \( \rightarrow \) \( \text{O}_2 - C'_{\text{CO}_2} = (1.01)C_{\text{CO}_2} \)

Figure 19 Continued.
21.6 \[ C'_{\text{e2}} = 1.01 \times C_{\text{e2}} \]

**Experimental**
- ○ \( H_2O \)
- □ \( O_2 \)

**Theoretical**
- \( C_{\text{e2}} \)
- \( \times \) \( H_2O - C_{\text{e2}}' = (1.01)C_{\text{e2}} \)
- \( O_2 - C_{\text{e2}}' = (1.01)C_{\text{e2}} \)

**Figure 9. Concluded.**
Figure 20. Species concentration profiles for a decrease in $C_2$ by one order of magnitude.

\[ 8.26 \quad C_2 = 0.1 \times C_{92} \]
$15.4 \quad C_2 = .1 \times C_{92}$

**Experimental**
- $H_2O$
- $O_2$

**Theoretical**
- $C_2 = C_{92}$
- $H_2O \quad C_2 = (0.1)C_{92}$
- $O_2 \quad C_2 = (0.1)C_{92}$

**Figure 28** Continued.
Figure 26 Concluded.
Theoretical \[ t_{1/2} = 0.03 \]  
\[ t_{1/2} = 0.03 \]  
\[ t_{1/2} = 0.03 \]

Figure 21 Species concentration profiles for an increase in initial \( t_{1/2} \) profile by one order of magnitude.
$15.4 \quad t^{12} = 0.03$

**Experimental**
- ○ $H_2O$
- □ $O_2$

**Theoretical**
- $t^{12} = 0.003$
- $H_2O - t^{12} = 0.03$
- $O_2 - t^{12} = 0.03$

Figure 11 Continued.
21.6 $t^{1/2} = 0.03$

**Experimental**
- $\bigcirc$ H$_2$O
- $\square$ O$_2$

**Theoretical**
- $\bigcirc$ H$_2$O - $t^{1/2} = 0.003$
- $\times$ H$_2$O - $t^{1/2} = 0.03$
- $\bigcirc$ O$_2$ - $t^{1/2} = 0.03$

Figure 21 Concluded.
Figure 22. Experimental initial k-ε profiles.
Figure 23 Species concentration profiles for input experimental k-6 profiles ($\phi = 1.0$).
Figure 23 Continued.
Figure 23 Concluded.
Figure 14. Species concentration profiles for input experimental $k$-$\varepsilon$ profiles ($\phi = 0.1$).
Figure 24 Continued.
Theoretical profiles

Experimental
- O H₂O
- O₂

Theoretical
- Calculated profiles
- X H₂O - Input profiles
- + O₂ - Input profiles

Figure 2A. Concluded.
Figure 25 Species concentration profiles for input experimental k-ε profiles with φ = 1.0 and φ = 0.1.
Schematics of the Cookson Experiments

Figure 27
FIGURE 2B MINIMUM HYDROGEN FOR SUSTAINED COMBUSTION OF KEROSENE. (COOKSON)
The Effect of Fuel Hydrogen Concentration on the Ignition Delay Times of Propane-Hydrogen Fuels at One Atmosphere and Equivalence Ratio of One

$T_0 = 1000\,\text{K}$

$p = 1\,\text{atm}$

$\phi = 1.0$

Original page is of poor quality
FIG. 31

THE EFFECT OF HYDROGEN CONCENTRATION ON THE IGNITION DELAY TIME AT $\phi = 0.2$

Ignition Delay Time, seconds

5 × 10^{-2}

10^{-2}

10^{-3}

10^{-4}

10^{-5}

10^{-6}

Reciprocal Initial Temperature, $K^{-1} × 10^4$

5

6

7

8

9

10

70 C_3H_8/30 H_2

$\phi = 0.2$

$\rho = 1 \cdot 10^{-6}$
REACTION TIMES FOR STOICHIOMETRIC
PROPANE-HYDROGEN FUELS AT ONE
ATMOSPHERE

\[ \text{Reciprocal Initial Temperature, } K^{-1} \times 10^4 \]

\[ \phi = 1 \]
\[ p = 1 \text{ atm} \]
BLOWOUT LIMIT CORRELATIONS FOR PROPAPE-HYDROGEN FUELS AT ONE ATMOSPHERE AND G00K INITIAL TEMPERATURE

FIG. 34

Equivalence Ratio

0.1 1.0 3.0 10.0 30.0
Mass Flow Rate per Unit Volume

70% C_3 H_8 / 90% H_2 (by vol.)

O 90/10
O 80/20
O 60/40
The Effect of Fuel Hydrogen Concentration on the Blown Ratio

Initial Temperature Limit at Unit Equivalence Ratio of Atmosphere and 600K

Fig. 35
APPENDIX A. The Ranges of Applicability of Chemical Kinetic Models of Hydrogen-Air Combustion

A.1. Introduction

A.1.1 Motivation for the Present Study

The study of hydrogen-fueled supersonic combustion ramjets (scramjets) has been a major part of the research programs of the Hypersonic Propulsion Branch at the Langley Research Center. Scramjets take advantage of high flight Mach numbers (greater than 5) to achieve propulsion efficiencies greater than that of ramjet engines. Propulsion efficiency is a measure of the thrust energy out divided by the combustion energy in. Due to the great flight speeds of scramjets (Mach numbers range from 4 to 7), there are time constraints for combustion within a combustor of reasonable size.

As a result of these time constraints, minimal ignition delay times are desirable. Figure 1 shows a plot of the log of the ignition time versus initial temperature for various fuels. Hydrogen fuel has a very low ignition delay time relative to hydrocarbon fuels. It is this low ignition delay time which makes hydrogen a valuable fuel for scramjets. It becomes a necessity, therefore, to study in detail hydrogen-air combustion.

Most studies, thus far, have been experimental investigations. This dependence on empirical results is due to the complexity of the flow around fuel injectors with three-dimensional geometries, which are not easily treated analytically. Numerical solutions have generally been
restricted to two- or three-dimensional parabolic flow with oversimplified chemistry models of the $\text{H}_2$-air system. These numerical solution schemes are applicable only in the parabolic flow region well downstream of the disturbance.

![Figure 1: Ignition delay time vs. initial temperature](image)

FIGURE 1 IGNITION DELAY TIME VS. INITIAL TEMPERATURE

caused by the transverse fuel injection used by scramjet combustors in order to achieve rapid mixing and reaction. A priori knowledge of the extent of fuel mixing, ignition, and reaction is required to initiate calculations.

In spite of the difficulty in obtaining quantitative information, a sufficient data base has been established to define a scramjet engine concept and to permit fabrication of subscale engine models with integrated inlet, combustor, and nozzle components. The current scramjet is designed to operate...
at stagnation temperatures between 900°F and 2200°F,
which correspond to a flight Mach number range from 4 to
7. In ground tests of subscale engine models, however,
problems were encountered in obtaining ignition and
sustaining reaction at test conditions where ignition and
sustained reaction were expected.

The need to better understand the chemical mechanism
of the ignition and reaction of H₂-air mixtures at
conditions typical of a scramjet combustor has led to many
analytical studies. Most of the analytical studies at the
Langley Research Center, as well as this present study,
use a computer program (references 4 and 5) to solve
flowing, chemical-kinetic, isobaric, stream-tube problems
involving many chemical species. It is known that the
computational time requirements for any computer program
employing detailed chemical kinetics is proportional to
the number of species and reactions being treated. It is
the purpose of this effort to develop a method which
reduces these numbers in the study of H₂-air combustion
and at the same time preserve the correct physical-
chemical behavior. The result of this work will be to
reduce computer time and computer storage requirements.
Methods for reducing computational times are presented in
section two of this work.
A.2. Methods to Reduce Computational Times

The method proposed in this work to reduce computer run times is an extension and refinement of the method proposed by Chintz in reference 9. This method involves the tracking of a "trigger" species which is used to determine whether the flow is in an "ignition" mode or in a "combustion" mode. Ignition and combustion are defined in the classical sense. Ignition delay time is taken to be the time required for the temperature increase to reach five percent of the overall temperature increase:

\[ \text{Tig} = T_0 + 0.05(T_{eq} - T_0) \]  

(1)

When the temperature is less than the ignition temperature, the flow is in the "ignition" mode. Otherwise, the flow is in the "combustion" mode. (See Figure 2).

It will be shown that while an extensive chemical package is needed to describe the "ignition" mode, a smaller package is sufficient to deal with the "combustion" mode. Evans and Schexnayder\textsuperscript{10} concluded that a 25-reaction scheme involving 12 species (designated 25(12) herein), was required to describe "ignition" processes, while an 8(7) sufficed to deal with the "combustion" mode. The work
discussed herein describes efforts to minimize the computational time requirements once the "combustion" process is initiated (i.e. a smaller package or one that takes less computational time than the 8(7) system might suffice to deal with the "combustion" mode).

A 37(13) system is used as the "test" chemical-kinetic package (Table 1). Any smaller system is "tested" against the 37(13) system to see if the correct physical chemical behavior is preserved. This is accomplished in the following way:

First, the full 37(13) system is run from the initial state to equilibrium. Next, the test system is run from the point where the "combustion" in the 37(13) system initiates. The test system's initial values (pressure, temperature, equivalence ratio and chemical composition) are that of the 37(13) system's values at the point of "combustion". Temperature-time profiles and chemical behavior are compared (see Figure 3). If the difference between the 37(13) system and the test system is within an acceptable
range, the test system will be sufficient to deal with the "combustion" process.

Graphs of the mass fraction of the trigger species versus time and ignition temperature versus mass fraction of the trigger species at ignition are made in accordance with reference 9. The cases from the 37(13) system serve as the data base.

An 8(7) system (Table 2) is tested to confirm the results of Evans and Schexnayder.10. The 8(7) system is run from the initial state to equilibrium to determine if there are any initial conditions where the 8(7) system describes the entire process (i.e. both ignition and combustion).

In order to further reduce computational times, a 2(5) global model (Table 3) developed by Rogers and Chinitz in reference 11 is tested. The method proposed in that paper must be altered to include the effects of pressure. The Arrhenius equation will have the form:

$$k_{fi} = A_i(\varphi, \rho)T^{Ni}\exp(-E_i/RT) \quad (2)$$

The values of the parameters may be different from those in reference 11 because they are fixed
arbitrarily so that the 2(5) system describes the 37(13) system.

Lastly, the so-called "partial equilibrium" assumption of reference 12 is examined in connection with an 8(7) system in an effort to minimize computational times. The 8(7) mechanism consists of the following reactions:

\[
\begin{align*}
H_2O + O &= OH + OH \\
H_2O + H &= OH + H_2 \\
O_2 + H &= OH + O \\
H_2 + O &= OH + H
\end{align*}
\]

\[
\begin{align*}
H_2O + M &= H + OH \\
H_2 + M &= H + H + M \\
O_2 + M &= O + O + M \\
OH + M &= O + H + M
\end{align*}
\]

During the ignition delay period, all eight reactions play a role; however, in the "combustion" mode, the bimolecular shuffle reactions occur so rapidly in both directions that under some conditions they may be basically in equilibrium. The "partial equilibrium" assumption takes reactions (i)-(iv) to be
in equilibrium (infinitely fast reaction rates), while reactions (v)-(viii) continue to occur at a finite rate.

Infinitely fast reaction rates can be approximated by assigning very large values to the preexponential factor in the forward reaction rate constant (e.g. $10^{50}$). The backward reaction rate constant is then very large to satisfy $K = k_f / k_b$. The values of the preexponential factors of reaction (i) to (iv) are determined so that the system describes the "combustion" process.

To obtain these results, 45 one-dimensional, constant pressure, H₂-air computations were performed in the ranges:

\[ 0.5 \leq \phi \leq 1.5 \]
\[ 850 \leq T \leq 1200 \]
\[ 0.5 \leq p \leq 1.0 \]

where $\phi$ is the equivalence ratio, $T$ is the temperature in °K, and $p$ is the pressure in atmospheres. A listing of these individual cases is given in Table 4. Comparison of the computational times are then made and recommendations are discussed.
A.3. Results and Discussion

A.3.1 Summary

All 45 cases in Table 4, as well as the nine cases marked with an exclamation point, were performed on the 37(13) and 8(7) systems, respectively, in both the ignition and combustion modes. In addition, the 15 cases marked with an asterisk were tested on the 8(7) system, the global model, and the "partial equilibrium" model starting at the ignition point. These 15 cases were selected to encompass the entire range of interest.

What follows are the results of these test cases. First, the results from the 37(13) system including the "trigger" concept are presented. Next, the 8(7) system in its entirety and in the combustion mode alone is discussed. Global and partial equilibrium results in the combustion mode follow. Lastly, temperature-time profiles and computational times are compared for all four systems.

A.3.2 The 37(13) System

A.3.2.1 Temperature-Time Profiles

45 cases were run to serve as a basis of comparison for the test systems described
previously. Sufficient data points were obtained for all cases and each case yielded a characteristic S-shaped curve for the temperature versus time profile. For each case, it was necessary to obtain a data point at the ignition temperature, defined in the classical sense as 5% of the temperature rise from the initial state to the final equilibrium state, in order to set initial conditions for the test cases. Problems arose, however, when trying to pinpoint the ignition temperature. As a result, satisfactory ignition conditions could not be calculated for nine cases due to time constraints. These cases were omitted from data base and the remaining 36 cases are given in Table 5 with corresponding ignition temperature, time, and mass fraction of OH.

Tables 4 and 5 show that the ignition delay time is a function of initial temperature, pressure, and equivalence ratio. This function is complicated since it is inversely proportional to initial temperature, nearly independent of equivalence ratio except at low temperatures (850°K and 900°K), proportional to pressure at temperatures of 850°K and 900°K, and is inversely proportional to pressure at temperatures greater than or equal to 1000°K in the ranges studied. These results are similar to the
results reported by Rogers and Schexnayder (reference 2) in their extensive study which included 60 reactions and 20 species. Since the results are similar to those of Rogers and Schexnayder, the 37(13) system serves as a good basis for the remaining systems to be compared with.

A.3.2.2 The "Trigger" Concept

In order to test systems in the "combustion" mode it is necessary to determine when the transition to the "combustion" mode takes place. It was to this end that the concept of a "trigger" species was first proposed by Chinitz in reference 9. The concept involves the tracking of the mass fraction of one of the species. When its value becomes greater than a preset value, the system is said to be in the "combustion" mode.

There are several requirements for a species to serve as an effective trigger from ignition to combustion. The concentration of an effective trigger species must be single-valued (the value of the mass fraction at the ignition point is not again obtained) and undergo a sufficiently large change during the ignition process that its crossover into the combustion mode is unmistakeable. It was
believed that the hydroxyl radical, OH, would serve that purpose.

Results here confirm the work of reference 9. Table 6 and Figure 4 show how the mass fraction of OH varies with time for a representative case (case number 23). It is easily seen that OH has all the requirements to serve as the trigger species. The ignition point is well-defined and is single-valued.

It is of interest to estimate the trigger point for all values of equivalence ratio, pressure, and temperature studied. A plot of the ignition temperature vs the mass fraction for OH at the ignition point is shown in Figure 5. A least-squares linear fit of \( \ln T_{ig} \) is \( Y_{OH_{ig}} \) was determined to be

\[
\ln T_{ig} = 0.085 \ln Y_{OH_{ig}} + 7.65 \quad (3)
\]

This linear relationship predicts values within a factor of 2.5 (generally well within) for all pressures and equivalence ratios examined. This provides a large advantage over the non-linear relationship developed in reference 9 which was limited to an equivalence ratio of 1. With the selection of a trigger species whose value can be estimated, the test systems can now be examined.
A.3.3 The 8(7) System

A.3.3.1 Ignition and Combustion

The 8(7) system was run from time zero to equilibrium for the nine cases marked with an exclamation point in Table 4. Five of these cases (3, 6, 9, 12, 15) represent a set of constant pressure, constant equivalence ratio conditions with varying initial temperature. These were selected to determine how the initial temperature affects the ignition delay time when compared to the 37(13) system. The remaining four cases (13, 22, 25, 29) were arbitrarily picked in an attempt to see the effects of pressure and equivalence ratio.

Figures 6-10 show a comparison of temperature-time histories for the 37(13) and 8(7) systems for cases 3, 6, 9, 12, 15, respectively, in both the ignition and combustion modes. Figures 6 and 7 show the delay time to be more than an order of magnitude lower for the 8(7) system at low temperature. Figures 8-10, however, show that the delay time for the 8(7) system is greater than the 37(13) system at higher temperatures. It is therefore concluded that a transition takes place between an initial temperature of 900°K and 1000°K. (It should be noted that a ten percent mixture of hydrogen in air has an ignition temperature of between 893°K and 1020°K\textsuperscript{13}).
Figure 11 shows a plot of ignition delay time vs. $1000T_0^{-1}$ for the 8(7) and 37(13) systems at a pressure of one atmosphere and an equivalence ratio of 0.5. The two curves cross at a temperature of 970 K where their ignition delay times are equal.

These results imply that for any constant pressure and constant $\phi$ there is a $T_0$ for which the two systems have the same ignition delay time. At these conditions, the 8(7) system can describe the 37(13) system in both the ignition and combustion modes. It also might be expected that at temperatures away from this particular initial temperature, the 8(7) system might lag behind or proceed ahead of the 37(13) system when tested at ignition conditions in the "combustion" mode.

The remaining four cases all had initial temperatures at or above 1000 K. The results of these cases were again that the delay time was higher for the 8(7) system than that of the 37(13) system. No significant effects of pressure or equivalence ratio were found in these limited case. The 8(7) system was now tested in the "combustion" mode with a focus on the effect of initial temperature.
A.3.3.2 The "Combustion" Mode

The 15 cases marked with an asterisk in Table 4 were run for the 8(7) system in the "combustion mode." For all pressures, temperatures, and equivalence ratios studied, the 8(7) system had a higher equilibrium temperature than that of the 37(13) system. This result was expected in that the 8(7) system neglects nitrogen dissociation. It was also found that for all conditions, the 8(7) system lagged behind the 37(13) system after the ignition point then equaled and surpassed the 37(13) curve (see Figure 12). This result was contrary to the expectations of section 4.3.2.1. It was thought that for some initial temperatures the 8(7) system would reach equilibrium more quickly than the 37(13) system. This, however, was not the case.

The amount of lag, however, was found to be a function of the relationship between the actual ignition temperature and the ignition temperature calculated from equation 1 (designated herein as "classical" ignition temperature). If the ignition temperature used (and also all other initial conditions) was between the initial temperature and the classical ignition temperature, then the 8(7) system described the 37(13) system very well. If the temperature used was greater than the classical
ignition temperature, there was an appreciable lag (see Figure 12). Therefore, as the difference between the initial temperature and ignition temperature used increased, the lag became greater. No cases were found where the 8(7) system reached equilibrium more quickly than the 37(13) system. It is concluded, therefore, that ignition temperatures lower than the classical ignition temperature should be used. It was not determined what percent below five percent should be used.

A.3.4 The Global Model

The global model in Table 3 was first proposed in reference 11. The reaction rate constant for each reaction is adjusted so that the global model describes the 37(13) system in the combustion mode.

Before determining these adjusted rate constants, it was useful to observe how varying each rate constant affected the temperature-time profile. It was observed that increasing $k_f 4$ (the subscript 4 refers to the first reaction; subscript 5 refers to the second reaction using the notation in reference 11) slows down the reaction. This was referred to as a "lagging" system in the previous section. Increasing $k_f 5$ speeds up the reaction. The system is very sensitive to an increase or
decrease in $k_{f5}$ while being rather insensitive to
a variation in $k_{f4}$. It was decided to fix Ni and
Ei at the same values as in reference 10; namely,

\[
\begin{align*}
E_4 &= 4865 \text{ cal/mol} \\
N_4 &= -10 \\
E_5 &= 42,500 \text{ cal/mol} \\
N_5 &= -13
\end{align*}
\]

Due to the insensitivity of the system to
$k_{f4}$, it was also decided that $A_4$ could be
retained as in reference 11:

\[
A_4 (\phi) = (8.917\phi + 31.433)\phi - 28.950 \times 10^{47}, \text{ cm}^3/\text{mol-s.}
\]  \hspace{1cm} (4)

It should be noted that $A_4$ is a function of $\phi$ only.

The effect of pressure, which was neglected in
reference 11, was included in the determination of
$A_5$. As a result, $A_5$ is a function of pressure
and equivalence ratio and has the form $A_5 (\phi, P)$. The sensitivity of the system to $k_{f5}$ permitted
substantial changes to be made to the temperature
profiles produced by the global system.

Table 8 shows the values of $A_4$ and $A_5$
determined in order to describe the temperature-time
profile required. Figure 13 shows a plot of $A_4$
(which is independent of pressure) vs. $\phi$. Figure 14
shows plots of $A_5$ vs. pressure for various $\phi$'s.
As other authors have reported, large discrepancies exist between results for equivalence ratios less than or greater than one. The relationship between $A_5$ and pressure and equivalence ratio must be broken up into two equations for the ranges studied:

**For $\phi \geq 1$**

$$A_5 = \left( \frac{8.80 + 5.85/\phi - 3.67\phi - 4.80p - 3.20p/\phi + 2.00p\phi}{10^{6.4}} \right)$$

**For $\phi < 1$**

$$A_5 = 0.67\phi^{-1.5} p^{-0.7} (8.80 + 5.85/\phi - 4.80p - 3.20p/\phi + 2.00p\phi) x 10^{6.4}$$

These equations predict values of $A_5$ to within 10% for the $\phi$'s studied.

In all cases, the global mode lags behind the 37(13) system for a short period following ignition. Then the global model equals and exceeds the 37(13) system and reaches equilibrium (see Figure 15). The equilibrium temperature of the global model should be higher than that of the 37(13) system due to dissociation.

This again, however, is controlled by the ignition temperature used by the global model. If the ignition temperature used is greater than the classical ignition temperature, then the equilibrium temperature is less than the classical ignition temperature. If the ignition temperature is less
than the classical, then the equilibrium temperature reached is greater than the 37(13) system as expected.

Initial temperature also has an effect on the accuracy of the global model. As the initial temperature increases, the difference between the global model and the 37(13) system increases. No effect on pressure or equivalence ratio was found.

The global model predicts very well the temperature-time profile of the 37(13) system. These results, however, show the importance of the ignition temperature used. It is suggested that ignition temperature be a variable to be tested in future work. Nonetheless, the global model shows promise in accurately reproducing temperature-time profiles with a minimum of computational time and computer storage requirements.

A.3.5 The "Partial Equilibrium" Assumption

The "partial equilibrium" assumption is already in use in combustion analysis; however its applicability to this present study in the ranges of interest has not been determined. This assumption states that the bimolecular shuffling reactions occur
so rapidly that they are basically in equilibrium. If this is the case, the kinetics equations can be replaced by the algebraic laws of mass action as discussed in reference 12. Therefore, a number of partial differential equations are replaced by algebraic equations and the remaining partial differential equations are simplified. If appropriate criteria are specified as to when this assumption is applicable, computer running times will be greatly reduced.

It was to this end that an approximation to the "partial equilibrium" assumption was tried. Instead of incorporating algebraic equations into the computer program, infinitely fast reaction rates would be approximated by using extremely large numbers (on the order of $10^{50}$) for the preexponential factor. It was anticipated that this would sufficiently approximate the "partial equilibrium" assumption.

Substantial numerical difficulties arose in trying the approximation. First, the largest preexponential factors that could be used were of the order of $10^{24}$. Still, this was at least 5 orders of magnitude greater than any other preexponential factors. Next, the precision of the program needed to be upgraded two orders of magnitude by adjusting the
EMAX parameter. Finally, this approach produced erroneous results such as temperatures exceeding the equilibrium temperature when the Gear integration procedure was used. The slower Adams method was then used which gave results which were physically plausible. Due to these problems, running times were very long. This was not the main concern, however. Rather, the primary goal was to test the partial equilibrium assumption against the 37(13) system in the ranges studied.

Figures 16-30 show how the approximation to the "partial equilibrium" assumption compared to the 37(13) system. For low temperatures (850°K, 900°K) and for all cases with an equivalence ratio of 0.5, the partial equilibrium assumption reproduced the 37(13) curve well. In the remaining cases, however, the assumption did not approximate the profile accurately. The shape of the curve was not even preserved.

It appears that the approach to approximating the "partial equilibrium" assumption has a narrow region of validity; namely, low initial temperatures and low equivalence ratios. As stated before, no conclusion can be drawn about the assumption itself. It is suggested that future work include the
incorporation of algebraic laws of mass action into the computer program to more precisely determine the validity of the assumption.

A.3.6 Comparison of the Four Systems

A.3.6.1 Temperature-Time Profiles

Figures 16-30 show a comparison of the 37(13) systems for the 15 cases marked with an asterisk in Table 4. The ignition mode of all 15 cases is that of the 37(13) system which continues in the combustion mode. The test systems start at the ignition point and reach an equilibrium temperature different from the 37(13) system.

The lag of the 8(7) system is shown in these figures. The lag is more a function of the relationship between the ignition temperature used and the classical ignition temperature than the initial temperature, pressure, or equivalence ratio. The 8(7) system reproduces the temperature-time curve very well, as expected.

The initial lag and eventual higher equilibrium temperature of the global model is shown in all 15 cases. Initial temperature has a substantial effect on the accuracy of the global model. As the initial temperature increases, the temperature differences
between the global model and the 37(13) curves become large. These perhaps could be corrected by starting the global model at a different ignition temperature.

Partial equilibrium curves accurately reproduced the temperature-time curves for cases 2, 4, 9, 11, 15, 18, 20 and 31. This suggests that the partial equilibrium assumption might suffice at lower initial temperatures and equivalence ratios. The numerical difficulty in obtaining these results sheds uncertainty on this conclusion.

Figures 31 and 32 represent the results of using an ignition temperature higher than the classical one to cases 11 and 15, respectively. The 37(13) system is observed to lag behind the 37(13) system to a greater extent than usual. The global model, in fact, has an equilibrium temperature which is 200 degrees below normal. These figures reinforce the importance in the selection of initial conditions for any test system.

A.3.6.2 Computational Time

Results of the average computational times for the various systems are shown in Table 9. The global model has the lowest computational time as expected, with an average time
of approximately 1/8 that of the $37(13)$ system.

Excessive computational time of the partial equilibrium system was related to the relatively high reaction rate constants employed, the use of the Adams method rather than the Gear method, and the precision which was required to perform "partial equilibrium" calculations.
A.4. Conclusions and Recommendations

A.4.1 Summary

A data base, in the ranges of initial temperature between 850°K and 1200°K, pressure between 0.5 and 1.00 atmospheres and equivalence ratio between 0.5 and 1.5, was established with the 37(13) system serving as the basis. Due to the large amount of data produced, it is necessary to clarify these results. First, these results are summarized. Next, conclusions are enumerated. Finally, recommendations for future work are given.

The first part of the work was selecting a good trigger species. It was concluded that the hydroxyl radical, OH, serves as an effective trigger. The value of the mass fraction of OH at ignition can be estimated in the ranges studied within a factor of 2.5 (generally well within) according to equation 32.

Next the 8(7) system was studied in the "ignition" and "combustion" modes together and then in the "combustion" mode alone. It was concluded that a transition takes place between an initial temperature of 900°K and 1000°K (at a pressure of one atmosphere and an equivalence ratio of 0.5) where the ignition delay time of the 8(7) system switches from less than to
greater than the 37(13) system. It was further concluded that at an equivalence ratio of 0.5, pressure of 1.0 atmosphere and initial temperature of 970+30°K the 8(7) system can approximate the temperature-time profiles of the 37(13) system in both the ignition and combustion modes. It is recommended that future work include the investigation of conditions where the 8(7) system can describe the 37(13) system in both modes.

In the combustion mode alone, the 8(7) system reproduces the temperature-time profiles very well for all cases studied. The variables which were major factors in the accuracy of these profiles were the temperatures and species' concentrations used at ignition. If the temperature used was less than, but close to, the classical ignition temperature, then the profiles were very accurate. If, however, the temperature used was greater than the classical ignition temperatures, then the 8(7) system lagged behind the 37(13) system somewhat. Future work, should include the investigation of this phenomenon.

Next, the global model was studied in the combustion mode. Preexponential factors were determined with a rather large discrepancy between results for an equivalence ratio less than or greater than one. For all cases, the global model lags
behind the 37(13) system initially then exceeds it at equilibrium. Two variables had an effect on the accuracy of this model. First, the conditions used at ignition had a substantial effect on accuracy again. A lag was produced by using a temperature greater than the classical ignition temperature. Secondly, as the initial temperature increases the accuracy of the global model decreases.

Finally, an approximation to the "partial equilibrium" assumption was tried. Numerical difficulties arose in the execution of this approximation. Nonetheless, it was concluded that at low temperatures and low equivalence ratios the approach to approximating the partial equilibrium assumption is valid. No conclusions could be drawn about the "partial equilibrium" assumption itself.

Results for the various systems show the global model to have the lowest average computational time followed by the 8(7) system, the 37(13) system, and the "partial equilibrium" model. The two-step global model has a computational time of approximately 1/8 that of the 37(13) system.

A.4.2 Conclusions

The following conclusions can be drawn based on the results of this study:
The hydroxyl radical, OH, serves as an effective trigger species.

A transition takes place between 900°K and 1000°K (at a pressure of 1 atmosphere and an equivalence ratio of 0.5) where the ignition delay time of the 8(7) system switches from less than to greater than the 37(13) system.

At a pressure of one atmosphere, an equivalence ratio of 0.5 and an initial temperature of 970±30°K the 8(7) system can accurately approximate the temperature-time profile of the 37(13) system in both the ignition and combustion modes.

For all cases, the 8(7) system describes the 37(13) system very well in the combustion mode.

A lag is produced in the 8(7) and global models when using ignition temperatures greater than the classical ignition temperature.

The global model satisfactorily reproduces the temperature-time profiles of the 37(13) system with a minimum of computational time and computer storage requirements.

As the initial temperature increases, the accuracy of the global model decreases.

No conclusions can be drawn concerning the "partial equilibrium" assumption.

The approach to approximating the "partial equilibrium" assumption has a narrow region of validity; namely, low initial temperatures and low equivalence ratios.
(10) The global model can approximate the $37(13)$ system with an average of $1/8$ the computational time.

A.4.3 Recommendations

Based on the results of this study, the following recommendations are made for future work:

(1) Investigate conditions where the $8(7)$ system can accurately describe both the ignition and combustion processes.

(2) Study the effects of varying ignition conditions; namely, above and below the classical ignition temperature. This will show the relationship between ignition conditions and the lag of the system.

(3) Continue to investigate the two-step global model as it seems to have the most promise in reproducing temperature-time profiles with a minimum of computational time.

(4) Incorporate the algebraic laws of mass action into the computer program to more precisely determine the validity of the "partial equilibrium" assumption.
### Table 1 - 37(13) System

<table>
<thead>
<tr>
<th>REACTION NUMBER</th>
<th>M</th>
<th>+</th>
<th>1*O2</th>
<th>= 1*O</th>
<th>+</th>
<th>1*O</th>
<th>N2 is an inert species</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>M</td>
<td>+</td>
<td>1*H</td>
<td>+</td>
<td>1*H</td>
<td>+</td>
<td>1*OH</td>
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<td>1*H2</td>
<td>+</td>
<td>1*H</td>
<td>+</td>
<td>1*OH</td>
</tr>
<tr>
<td>3</td>
<td>M</td>
<td>+</td>
<td>1*H20</td>
<td>= 1*H</td>
<td>+</td>
<td>M</td>
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<td>+</td>
<td>1*O2</td>
<td>= 1*O</td>
<td>+</td>
<td>1*O</td>
<td></td>
</tr>
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<td>5</td>
<td>M</td>
<td>+</td>
<td>1*NO</td>
<td>+</td>
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<td>+</td>
<td>1*O</td>
<td></td>
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<td>+</td>
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<tr>
<td>13</td>
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<td>+</td>
<td>1*O</td>
<td>= 1*H</td>
<td>+</td>
<td>1*OH</td>
<td></td>
</tr>
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<td>14</td>
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<td>+</td>
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<td>= 1*O</td>
<td>+</td>
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<td></td>
</tr>
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<td>16</td>
<td>1*H2</td>
<td>+</td>
<td>1*O2</td>
<td>= 1*H</td>
<td>+</td>
<td>1*OH</td>
<td></td>
</tr>
<tr>
<td>17</td>
<td>1*H</td>
<td>+</td>
<td>1*H2</td>
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N₂ is an inert species
### TABLE 3 - 2 STEP GLOBAL MODEL

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\begin{align*}
  & H_2 + O_2 \rightarrow 2*OH & (1) \\
  & 2*OH + H_2 \rightarrow 2*H_2O & (2)
\end{align*}
\]

\(N_2\) is an inert species
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<td>$5.0E47$</td>
<td>$4.0E64$</td>
</tr>
<tr>
<td>41</td>
<td>$5.0E47$</td>
<td>$4.6E64$</td>
</tr>
<tr>
<td>43</td>
<td>$5.0E47$</td>
<td>$5.1E64$</td>
</tr>
</tbody>
</table>
### TABLE 9

**COMPARISON OF COMPUTATIONAL TIMES**

<table>
<thead>
<tr>
<th>SYSTEM</th>
<th>AVERAGE COMPUTATIONAL TIME (seconds)</th>
</tr>
</thead>
<tbody>
<tr>
<td>37(13)</td>
<td>87.5</td>
</tr>
<tr>
<td>8.(7)</td>
<td>12.3</td>
</tr>
<tr>
<td>Global</td>
<td>11.4</td>
</tr>
<tr>
<td>Partial</td>
<td>1258</td>
</tr>
</tbody>
</table>
REFERENCES
FOR APPENDIX A

1) Chinitz, W., Personal Communication, Cooper Union, New York, Apr., 1983.


FIGURE 2 DEFINING IGNITION AND COMBUSTION

FIGURE 3 COMPARISON OF TEMPERATURE-TIME PROFILES FOR 37(13) AND TEST SYSTEMS
FIGURE 4 MASS FRACTION OH VS TIME

FIGURE 5 IGNITION TEMPERATURE VS.
MASS FRACTION OH

\[ \ln T_K = 3.905 \ln Y_{OH} + 7.65 \]
**Figure 6** \( \phi = 0.5, P=1.0, T_0=850 \)

**Figure 7** \( \phi = 0.5, P=1.0, T_0=900 \)
FIGURE 10 $\phi=0.5$, $P=1.0$, $T_0=1200$

FIGURE 11 Ignition Time vs $1000 \cdot T_0$
FIGURE 12 COMPARISON OF 37(13) AND B(7) SYSTEMS IN COMBUSTION MODE

FIGURE 13 A₄ VERSUS $\phi$
FIGURE 15 COMPARISON OF 37 F3 AND GLOBAL SYSTEMS IN COMBUSTION MODE

FIGURE 14 A5 VS. F FOR VARIOUS $\phi$'S
FIGURE 18 \( \phi = 0.5, p = 1.00, T_0 = 1000 \)

FIGURE 19 \( \phi = 0.5, p = 0.75, T_0 = 1100 \)
**FIGURE 20** $\phi = 0.5, P = 1.0, T_c = 1200$

**FIGURE 21** $\phi = 1.0, P = 1.0, T_c = 850$
FIGURE 22 $\phi=1.0, p=0.75, T_e=500$

FIGURE 23 $\phi=1.0, p=0.75, T_e=1000$
FIGURE 24  $\phi=1.0, P=0.5, T_c=1100$

FIGURE 25  $\phi=1.0, P=0.5, T_c=1200$
FIGURE 26 \( \phi = 1.5, P = 0.5, T_c = 850 \)

FIGURE 27 \( \phi = 1.5, P = 1.0, T_c = 900 \)
FIGURE 30 \[ \phi = 1.5, P = 0.5, T_c = 1200 \]

FIGURE 31 LAG FOR CASE NO. 11
FIGURE 32 LAG FOR CASE NO. 15