ABSTRACT

Oxygen concentration and separation is an essential factor for air recycling in a CELSS. Furthermore, if the value of the plant assimilatory quotient is not coincident with that of the animal respiratory quotient, the recovery of O₂ from the concentrated CO₂ through chemical methods will become necessary to balance the gas contents in a CELSS. Therefore, oxygen concentration and separation equipment using Salcomine and O₂ recovery equipment, such as Sabatier and Bosch reactors, were experimentally developed and tested.

1. BASIC CONSIDERATION ON GAS RECYCLING

Fundamental functions of gas recycling in CELSS, as shown in Fig. 1.1, are to separate each component, such as O₂, CO₂, and N₂ gases, and to concentrate and store each gas in order to supply appropriate gas concentration for human beings, plants and algae.

2. OXYGEN RECOVERY FROM CONCENTRATED CARBON DIOXIDE

Many research and development efforts have been conducted for the recovery of oxygen from concentrated CO₂ by the catalytic hydrogenization process. The Bosch reaction utilizes an iron catalyst and produces carbon (C) and water (H₂O) with CO as an intermediate product /1/. Product water is further electrolyzed to recover O₂ for animal respiration and H₂ for subsequent hydrogenization. The Sabatier reaction utilizes a ruthenium (Ru) catalyst and produces methane and water /2/. An additional CH₄ cracking process providing C and H₂ is necessary to make the Sabatier reaction comparable to the Bosch reaction /3/.

Mitsubishi Heavy Industries, Ltd. (MHI) has been involved in CELSS research for several years /4/ under the direction of the National Aerospace Laboratory and conducted the experimental program to evaluate the basic characteristics and performance of the two oxygen recovery processes.

2.1 Once Through Tests

In order to determine the basic characteristics of each reaction, once through tests of CO₂ reaction with H₂ on the catalyst are performed. Fig. 2.1 shows the schematic of the test set up.

Bosch Reaction

Fig. 2.2 shows the reaction rate of CO₂ vs. reaction temperature in the first Bosch reaction, as well as the reaction rate of CO vs. temperature in the second Bosch reaction. The different nature of the two reaction rate curves suggests a phased process concept using two reactors operating at different temperature ranges to obtain higher performance in the Bosch CO₂ cracking process.
Sabatier Reaction  Fig. 2.3 shows the reaction rate of CO$_2$ vs mole ratio of H$_2$/CO$_2$ in the feed gas mixture of the first Sabatier reaction. Reaction rate of more than 99% is achieved with a little bit higher ratio (4.5) than the stoichiometric value. As for the second Sabatier reaction, reaction rate of CH$_4$ vs. time after start of the reaction is presented in Fig. 2.4. Very rapid degradation is observed for the catalytic reactions using Pt or Ni, while a steady and much higher conversion efficiency is observed in pyrolytic reaction on silica wool filler.

2.2 Recycle Test

A Recycle Test Apparatus was prepared based on the once through test results as shown in Fig. 2.5.

The supply rate of mixture gas is shown in Fig. 2.6. The amounts of processed CO$_2$ are 0.15 kg/day and 0.42 kg/day for Bosch and Sabatier, respectively. Energy required for the above processes were estimated and the Sabatier process showed several times less energy use than Bosch process. The character of deposit carbon in the Bosch process was a loose powder form, while a hard solid block was obtained from the Sabatier reactor. This fact has an important meaning for maintenance operations to periodically extract carbon on orbit.

Thus, the O$_2$ recovery system with Sabatier methane cracking shows a good possibility for application in CELSS and Space Station.
3. OXYGEN SEPARATION SYSTEM USING SALCOMINE

3.1 Functions of Oxygen Separation System

Kawasaki Heavy Industries Ltd. (KHI) has been involved in CELSS research under contract with the National Aerospace Laboratory (NAL) /5/. Fig. 3.1 shows the function of O₂ separation and concentration system, concerning the CO₂ and N₂ separation and concentration. In the Fig. 3.1, the inlet gas is mixture of O₂, CO₂ and N₂. At the CO₂ concentrator, CO₂ is separated and concentrated, and stored in the CO₂ gas bottle. At the O₂ concentrator, O₂ is separated and concentrated, and stored in the O₂ gas bottle. The residual is N₂ gas and this is stored in the N₂ gas bottle. Thus, inlet gas is separated into three (3) components, CO₂, O₂, and N₂ having high purities.

There are many methods for CO₂ separation and concentration. The studies reported here are focused on the O₂ separation and concentration system using Salcomine. Salcomine absorbs the O₂ under normal temperature (lower than about 40°C) and desorbs the O₂ under high temperature (higher than about 80°C). So, when the inlet gas is introduced into the Salcomine canister, the included O₂ is absorbed by Salcomine. After the Salcomine absorbs the O₂, the Salcomine canister is heated and Salcomine desorbs O₂ of high purity.

In the Fig. 3.1, there are three (3) canisters and each canister is operated in absorbing, desorbing and pre-cooling modes respectively. Thus, continuous O₂ separating operation can be carried out by the cyclic exchange of these operation modes.

3.2 SALCOMINE

Fig. 3.2 shows the structural formula of the Salcomine and its O₂ absorption and desorption. There are some studies on the application of Salcomine for O₂ concentration. Natsuda /6/ carried out the experimental study on the O₂ absorbing and desorbing characteristics and their variations under repeated reaction of the Salcomine, as a part of the development of Diffusive Atmosphere Control System (DACS, a kind of artificial gill), which extract O₂ from sea water dissolving oxygen.

The present study is based on these results and is designed to get data to determine the optimum O₂ separating and concentrating system for CELSS. In this study, O₂ absorbing and desorbing performance tests of Salcomine within the canister are carried out to get the data to design a compact, light weight and lower energy consuming system.

3.3 Oxygen Absorbing and Desorbing Test of SALCOMINE

3.3.1 Test of SALCOMINE Proper. In Fig. 3.3(a), oxygen containing nitrogen gas is introduced into the equipment and Salcomine absorbs the O₂, and increases its weight. And then the Salcomine is heated and the Salcomine decrease its weight desorbing O₂. From the weight change of the Salcomine, the O₂ absorbed and desorbed is measured. Test results are shown in Fig. 3.3(b). O₂ absorbing and desorbing capacity of Salcomine is expressed with the percentage of oxygen weight to the Salcomine weight (wt%). From Fig. 3.3(b), O₂ absorbing capacity is about 4.0 wt% for 20% of O₂ concentration.

3.3.2 Test of SALCOMINE within Canister. In Fig. 3.4, O₂ containing air is drawn by the air pump and introduced into the Salcomine canister. In the canister, O₂ in the air is absorbed by Salcomine.

Fig. 3.2 Oxygen Absorb and Desorb Reaction by SALCOMINE

Fig. 3.3 Characteristic Test of SALCOMINE Proper

(a) Test Equipment (b) Test Results

Page 69
Fig. 3.4 Diagram of Canister Contained SALCOMINE Performance Test

Fig. 3.5 Results of Canister Contained SALCOMINE Performance Test

And at the exhaust gas manifold, \( \text{O}_2 \) concentration is measured. During this time, the Salcomine canister is cooled. For the desorbing test, the canister is heated by hot water (90°C) and drawn by the vacuum pump. \( \text{O}_2 \) desorbed from the Salcomine is collected into the sampling bag and analysed. Fig. 3.5 shows the results of the test. Initially, \( \text{O}_2 \) concentration in the outlet air is almost 0% because all \( \text{O}_2 \) in the air is absorbed. After a few minutes, the break-through occurs and \( \text{O}_2 \) concentration rises, and reaches to 21% which is equal to that of the inlet air. From Fig. 3.6, absorbed \( \text{O}_2 \) is 675 CC and this corresponds to 2.7 wt% of \( \text{O}_2 \) absorbing capacity. On the start of \( \text{O}_2 \) desorbing test, the vacuum pump is operated to purge the air inside the test equipment. The canister is then heated. And is drawn by the vacuum pump. The drawn gas is measured 710 CC, and \( \text{O}_2 \) concentration in this gas is measured 91.4%. So, desorbed \( \text{O}_2 \) becomes 649 CC.

4. CONCLUSION

As a part of CELSS Gas Recycling System Development, preliminary experimental studies were conducted for the processes of \( \text{O}_2 \) recovery from concentrated \( \text{CO}_2 \) and of \( \text{O}_2 \) separation and concentration from exhaled gas of animal, plant etc. Sabatier process for the regenerable \( \text{O}_2 \) recovery from \( \text{CO}_2 \) shows a good possibility compared with Bosch process in its higher efficiency of \( \text{O}_2 \) recovery and easier handling of product carbon. From the \( \text{O}_2 \) separation and concentration tests using Salcomine, oxygen absorbing and desorbing characteristic data are obtained as the fundamental design data. The above results can be used in the planning of a complete and stable gas recycling system required in CELSS. We hope to continue our research and development effort for the expansion of the human frontier in Space.

REFERENCES


