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In the first year of the grant, potential energy surfaces (PES's) have been calculated for a number of reactions. The reactions currently under study are relevant to the design of the SCRAM jet engine for the NASP. Two hydrogen/oxygen reactions have been studied. The first reaction,

\[ H + O_2 \rightarrow HO_2 \rightarrow OH + O \]  

is the most important chain propagation reaction in combustion. Although there have been previous studies of reaction (1), they focused on the \( H + O_2 \) region of the surface and a detailed examination of the \( OH + O \) region was needed. This latter region of the surface is complex, because chemical bonding and long range electrostatic forces compete, leading to an \( OH - O \) minimum at large \( OO \) distances \(( \approx 6.0 a_0 )\) followed by a saddle point before the \( HOO \) well. Approximately 150 points have been obtained which characterize the minimum energy path (MEP) for the entire PES. Additional points will be computed to characterize the hydrogen atom transfer process. In collaboration with Ron Duchovic (Elorlet Institute), the complete set of points will be used to generate a global PES suitable for dynamical studies. A more detailed discussion of this work is given in ref. 2.

An additional hydrogen/oxygen reaction which has been studied is the reaction

\[ O(^1D) + H_2 \rightarrow H_2O^* \rightarrow OH + H. \]  

The reason for studying this system is that \( O(^1D) \) formed by photolysis of \( O_2 \) could be used as a means of initiating combustion at lower temperatures in the SCRAM jet engine. As discussed in the semiannual report for the first six months of this grant, both the vibrational energy distribution in the \( OH \) product and the \( OD/OH \) ratio when \( O(^1D) \) is reacted with HD seem to depend critically on the long range behavior of the potential. The new, more accurate, potential energy surface may explain discrepancies between
previous dynamical studies and experiment. A more detailed discussion of this work is given in ref. 3.

Another important group of reactions for design of the SCRAM jet are three body recombination reactions. As a start in this area, calculations were carried out for the reactions

\[ H + N_2 \rightarrow HN_2 \] (3)

and

\[ H + HN_2 \rightarrow H_2 + N_2 \] (4)

since the combination of eqns (3)-(4) constitutes a mechanism for H atom recombination with N\(_2\) acting as a third body. HN\(_2\) is also a very important intermediate in Thermal DeNO\(_x\) processes. Specifically, addition of NH\(_3\) leads to substantial NO\(_x\) reduction in hydrocarbon combustion. The key step here is thought to be

\[ NH_2 + NO \rightarrow \text{products} \]

where the expected products are HN\(_2\) + OH and N\(_2\) + H\(_2\)O. The observed branching ratio varies from study to study and while there is no direct evidence that HN\(_2\) is formed, OH is observed but H atoms are not. The presence of an HN\(_2\) intermediate is also implicated indirectly by modeling studies which require chain propagation steps producing OH in order to explain the DeNO\(_x\) process.

A key factor, currently undetermined, is the lifetime of the HN\(_2\) species. The potential for dissociation of HN\(_2\) has a shallow well \( \approx 3 \) kcal/mole above the H + N\(_2\) asymptote and separated from that asymptote by a barrier. The dissociation process is primarily controlled by quantum mechanical tunneling. Enough points on the HN\(_2\) dissociation potential have been obtained to characterize the MEP. Approximate calculations (one-dimensional tunneling through an Eckart barrier) have been carried out to determine
the rate of dissociation via tunneling. Additional electronic structure calculations with larger basis sets were carried out in collaboration with Celeste Rohlfing (Sandia National Laboratories, Livermore) to assess the accuracy of the computed potential. The larger basis set calculations indicate that the PES is converged to \( \approx \pm 1.0 \text{ kcal/mol} \). Based on this and the tunneling calculations the lifetime of \( \text{HN}_2 \) is estimated to be less than \( 10^{-8} \text{ sec} \). This is much shorter than values of about \( 10^{-4} \text{ sec} \) which would fit current models of thermal de-NO\(_2\) processes. This result suggests that there may be other important processes not included in the current models. A more detailed discussion of this work will be found in ref. 4.

Part of the potential for reaction (4) is shown in Fig. 1. Here the geometry of the \( \text{HN}_2 \) molecule is fixed at close to the computed equilibrium geometry and the distance between the second H and the center of mass of the \( \text{N}_2 \) as well as the polar angle are varied. From Fig. 1 it is seen that there are three channels which lead to stable minima (cis and trans \( \text{HNNH} \) and \( \text{H}_2\text{NN} \)) and a fourth abstraction channel which leads to \( \text{N}_2 + \text{H}_2 \). These results indicate that \( \text{HN}_2 \) is a very reactive species which is consistent with the short computed lifetime. At the current time it appears that the lifetime of \( \text{HN}_2 \) is too short for H atom abstraction from this species to contribute significantly to the \( \text{H} + \text{H} + \text{N}_2 \) three body recombination process; i.e. the dominant process may involve a different mechanism such as stabilization of a metastable \( \text{H}_2 \) species. Study of the latter process would require regions of the surface which were not considered in the present study; although, the \( \text{HN}_2 \) potential discussed above would be a good starting point.

\( \text{NO} \) will be an important species in the SCRAM jet engine. Since H atoms will also be present the potential surfaces for \( \text{HNO} \) are important. Currently studies are underway to characterize the lowest three surfaces (\( ^1\text{A}', \ 3\text{A}'', \text{and} \ 3\text{A}'' \)) of \( \text{HNO} \). Fig. 2 shows the potential for H atom addition to the N end of NO. The lowest surface, which involves addition to the radical orbital of NO, has no barrier, while the other two surfaces, which involve addition to the \( \pi \) bond of NO, have barriers. The \( \text{HNO} \) region of these three sur-
Fig. 1. The potential for the interaction of an H atom with HN$_2$. The HN$_2$ geometry is fixed at near the computed equilibrium geometry and the energy is shown as a function of i) the distance between H and the center of mass of N$_2$ and ii) the polar angle ($\theta$).

H + HN$_2$

faces has been characterized. The computed geometries for the $^1A'$ ground state and $^1A''$ excited state as well as the computed excitation energy are in good agreement with experiment. Calculations are also underway to examine the NOH region of the same three surfaces. These results will be combined with calculations in collaboration with
Celeste Rohlfing and Carl Melius at Sandia National Laboratory, Livermore on the O + NH and N + OH regions of the potential to produce a global potential for HNO.

Fig. 2. The energy as a function of $r_{NH}$ along the minimum energy path for H atom addition to the N end of NO for the lowest $^1A'$, $^3A''$, and $^1A''$ surfaces of HNO.

In the next year a number of projects will be completed. As mentioned above, additional points will be obtained for reaction (1). Studies on HNO will also be completed in collaboration with scientists at Sandia National Laboratory, Livermore. In collaboration with Ron Duchovic (Eloret), global representations of these potential surfaces will be constructed, and dynamical studies will be carried out using both statistical methods (RRKM theory and transition state theory) and classical trajectories. Further, quantum dynamical studies will be completed in collaboration with Dave Schwenke (Eloret). Finally, new studies using the improved O($^1D$) + H$_2$ potential energy surface will be carried out by George Schatz (Northwestern University).
Current Publications


Theoretical Characterization of the Minimum Energy Path for Hydrogen Atom Addition to N₂
Implications for the Unimolecular Lifetime of HN₂

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Abstract

The minimum energy path (MEP) for the addition of a hydrogen atom to N₂ is characterized in CASSCF/CCI calculations using the [4s3p2d1f/3s2p1d] basis set, with additional single point calculations at the stationary points of the potential energy surface using the [5s4p3d2f/4s3p2d] basis set. These calculations represent the most extensive set of ab initio calculations completed to date, yielding a zero-point corrected barrier for HN₂ dissociation of ≈ 8.5 kcal mol⁻¹. The lifetime of the HN₂ species is estimated from the calculated geometries and energetics using both conventional Transition State Theory and a method which utilizes an Eckart barrier to compute one-dimensional quantum mechanical tunneling effects. This study concludes that the lifetime of the HN₂ species is very short, greatly limiting its role in both termolecular recombination reactions and combustion processes.

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I. Introduction

The proposed National Aerospace Plane will be powered by an air-breathing, hydrogen-burning, supersonic ramjet (SCRAMjet) engine. The supersonic flow in the combustor and nozzle regions of the SCRAMjet will allow only very short residence times for the reactants, and consequently, accurate knowledge of the finite rates of reaction is critical to design of the engine. The rates of recombination processes are of special interest since these reactions can lead to considerable heat evolution and their rates, under conditions of high temperature and supersonic flow, are not well-known experimentally. One such reaction which is expected to be important under these conditions is the recombination of two hydrogen atoms in the presence of \( N_2 \):

\[
H + H + N_2 \rightarrow H_2 + N_2
\]  

The various mechanisms proposed to describe three-body recombination processes involve the formation of a metastable species. For reaction (1), the expected metastable species is either \( H_2^* \) (vibrationally-rotationally excited \( H_2 \)), which can be stabilized by subsequent collisions with \( N_2 \), or an \( HN_2 \) species, which can then react with a hydrogen atom to yield bound \( H_2 \):

\[
H + HN_2 \rightarrow H_2 + N_2
\]  

Calculations characterizing reaction (2) will be the subject of a future publication. We simply note here that these calculations indicate there is no barrier to reaction (2). However, in addition to the region of the potential energy surface which controls the abstraction process, there are three additional regions of the surface which lead
to stable \( H_2N_2 \) species (cis \( HNNH \), trans \( HNNH \), and \( H_2NN \)). While the \( HN_2 \) species has not been observed experimentally, indirect evidence for its existence has been obtained from experimental studies of the reaction:

\[
NH_2 + NO \rightarrow \text{products}
\]  

This reaction constitutes one possible step in the complicated mechanisms which describe the fundamental chemistry of nitric oxide both in the troposphere\(^1\) and in combustion\(^2-5\) processes. As a result, it has been studied extensively with a variety of experimental techniques over a range of temperatures and pressures. Early studies\(^6-15\) suggested that the major pathway is the reaction:

\[
NH_2 + NO \rightarrow N_2 + H_2O
\]  

Vibrationally excited \( H_2O \) and small amounts of the \( NH_2NO \) complex were observed along with a negative temperature dependence of the rate coefficient. Gehring et al.\(^11\) did entertain the possibility of an alternate pathway:

\[
NH_2 + NO \rightarrow N_2 + OH + H
\]  

However, a careful search using ESR techniques failed to identify any measurable quantity of hydrogen atoms.

In subsequent studies\(^16-29\), investigators considered a reaction mechanism which includes three possible product channels:

\[
NH_2 + NO \rightarrow N_2 + H_2O
\]  

\[
NH_2 + NO \rightarrow N_2 + OH + H
\]  

\[
NH_2 + NO \rightarrow N_2 + H_2O
\]  

\[
NH_2 + NO \rightarrow N_2 + OH + H
\]  

\[
NH_2 + NO \rightarrow N_2 + H_2O
\]
Each of these experiments attempted to determine the rate coefficient of the \( \text{NH}_2 + \text{NO} \) reaction as well as the major pathway of this reaction. The results of the studies are far from consistent. Stief et al. did not observe OH radicals, suggesting that the branching ratio for the formation of OH radicals is substantially less than 0.22 at 298 K. Further, both ESR techniques and Lyman-alpha fluorescence could not detect the production of hydrogen atoms. Hall et al., using infrared absorption at 298 K, reported that the OH channels account for only 13 ± 2% of the reaction, while Dolson attributed 15% of the reaction to OH radical production at 300 K. In contrast, Silver and Kolb determined that approximately 40% of the \( \text{NH}_2 + \text{NO} \) reaction produced OH radicals over a temperature range of 294 - 1215 K, although, at room temperature, they did not observe hydrogen atoms. Andresen et al., operating an isothermal flow reactor at temperatures ranging from 300 K to 1150 K, reported that the OH branching ratio is greater than 0.65 but that hydrogen atoms were not detected for temperatures from 300 K to 900 K. Finally, Kimball-Linne and Hanson observed a branching ratio for OH radical formation of 0.48 at 1050 K and a ratio of greater than 0.80 at 1400 K.

Despite the wide variation in the reported branching ratios for OH radical formation, the experimental results consistently indicate, over a range of temperatures, that hydrogen atoms are not produced. In light of equation (6), the observation of OH radicals without the production of hydrogen atoms indirectly argues for the existence of the rather exotic species, \( \text{HN}_2 \). However, none of the experimental studies have directly detected the presence of \( \text{HN}_2 \). Additional arguments for the existence of the \( \text{HN}_2 \) species have been made by researchers constructing kinetic
models of combustion processes\textsuperscript{3,4}. These models include several reactions of HN\textsubscript{2} as critical steps and require that this species possess a lifetime which is long enough to allow its participation in these reactions.

There have been a number of previous \textit{ab initio} studies\textsuperscript{30--35} of the reaction:

\[ HN_2 \rightarrow H + N_2 \] (7)

characterizing the geometry and energetics of the reactant, transition state, and products. The most recent of these earlier studies, the work of Curtiss et al.\textsuperscript{35}, concisely summarizes the methodologies and basis sets used in these investigations. At the SCF level of theory, the dissociation energy of HN\textsubscript{2} to H + N\textsubscript{2} was determined to have a value varying from -9.6 kcal mol\textsuperscript{-1} to +13.6 kcal mol\textsuperscript{-1} with a barrier to dissociation (before zero-point energy corrections) of 22.6 kcal mol\textsuperscript{-1}. The first two studies which treated electron correlation failed to resolve these early discrepancies. Vasudevan et al.\textsuperscript{31} estimated the dissociation energy of HN\textsubscript{2} to H + N\textsubscript{2} at 3.8 - 4.4 kcal mol\textsuperscript{-1} with an barrier height of 10 - 15 kcal mol\textsuperscript{-1}. Casewit and Goddard\textsuperscript{34} concluded that the dissociation energy is -20.0 kcal mol\textsuperscript{-1} but did not reinvestigate the barrier height. In their study, Curtiss et al.\textsuperscript{35} utilized third-order and fourth-order M\o ller-Plesset perturbation theory (MP3/MP4) to account for electron correlation effects and a double zeta basis set which included polarization functions (6-31G**) to compute a dissociation energy of -14.4 kcal mol\textsuperscript{-1} and a barrier of 10.5 kcal mol\textsuperscript{-1}. Using their calculated geometries and energetics in conjunction with Conventional Transition State Theory (including an estimate of quantal tunneling), Curtiss et al. then determined an approximate lifetime for the HN\textsubscript{2} radical.

While M\o ller-Plesset perturbation theory techniques work well in regions which
are well-described by single configuration SCF wavefunctions, difficulties do arise when bonds are stretched and when multiply-bonded systems such as N$_2$ are involved. Given the possible importance of the HN$_2$ species in a number of chemical systems, we decided to carry out an extensive calculation of the potential energy surface for the addition of hydrogen atom to N$_2$, using Complete Active Space/externally Contracted CI wavefunctions (CASSCF/CCI)$^{36}$ along with good quality Atomic Natural Orbital (ANO) basis sets$^{37}$. We have utilized these new calculations to study the lifetime of the HN$_2$ radical, using both Conventional Transition State Theory$^{38}$ and the formalism of W. H. Miller$^{39}$ which utilizes an Eckart Barrier to estimate one-dimensional quantum mechanical tunneling.

The \textit{ab initio} computational method is discussed in Section II. with the results of these calculations presented in Section III. Section IV. then discusses the dynamical methods used in this study along with an estimate of the lifetime of the HN$_2$ species. Finally, Section V. summarizes our conclusions.

II. Ab Initio Computational Details

Using a localized orbital description, the ground state of HN$_2$ is represented by equation (8):

\begin{equation}
\text{In equation (8), one NN } \pi \text{ bond has been broken and an NH } \sigma \text{ bond has been formed. Since the two bonds (one formed and one broken) are of about equal strength, a significant barrier is expected to occur in the entrance channel. Consequently, the}
\end{equation}
potential energy surface for hydrogen atom addition to \( N_2 \) has two stationary points, a saddle point for \( R_{NH} \approx 2.7 \ a_0 \) and a minimum for \( R_{NH} \approx 2.0 \ a_0 \).

ANO basis sets\(^3\) are used in the present study. Since these basis sets are optimal for describing the atomic correlation effects, they have very small basis set superposition errors, but they are sufficiently flexible to be used in molecular calculations at both the SCF and CI levels. The N and H basis sets are \((13s8p6d4f)/[4s3p2d1f]\) and \((8s6p4d)/[3s2p1d]\), respectively, and are described in detail in reference 37. While these basis sets were expected to describe the molecular potential energy curves accurately, additional calculations were completed with larger N, \((13s8p6d4f)/[5s4p3d2f]\), and H, \((8s6p4d)/[4s3p2d]\), basis sets, in which the additional contracted functions correspond to higher energy atomic natural orbitals. These additional calculations were carried out at both the stationary point geometries determined with the smaller basis sets, and the \( H + N_2 \) geometry.

Table I. describes the qualitative character of the CASSCF orbitals for both \( H + N_2 \) and \( HN_2 \). The 1a' and 2a' orbitals correspond to the N 1s orbitals, which were not correlated in these calculations, while the 3a' and 4a' orbitals correspond to the N 2s orbitals. The 5a' and 9a' orbitals correspond to the \( N_2 \) \( \sigma \) bond pair. At large \( R_{NH} \), the 6a' orbital corresponds to the \( N_2 \ 1\pi_u \) orbital, the 7a' orbital to the H 1s orbital, and the 8a' orbital to the \( N_2 \ 1\pi_g \) orbital. At shorter \( R_{NH} \), there is considerable mixing of the orbitals which are appropriate in the limit of large \( R_{NH} \). Near the \( HN_2 \) minimum, the 6a' orbital corresponds to the \( N_2 \ 1\pi_u + H \ 1s \), the 7a' orbital to \( N_2 \ 1\pi_g + H \ 1s \), and the 8a' orbital to \( N_2 \ 1\pi_g - H \ 1s \).

Orbitals 4a' through 9a', and the 1a'' and 2a'' orbitals were active in the CASSCF calculations. This CASSCF treatment in six active a' and two active a'' orbitals is denoted as a (62) CASSCF. The inclusion of the 4a' orbital in the active space was not conceptually necessary based on equation (8), but it was found that the larger
active space led to better CASSCF convergence, even though the CASSCF wavefunction of this orbital contained no configurations with CI coefficients greater than 0.05. In the CCI calculations, all but the N 1s electrons (the 1a' and 2a' orbitals) were correlated. The CCI list of reference configurations was obtained by selecting the unique configurations from the H + N₂ region, the saddle point region, and the HN₂ minimum region whose CI coefficients in the CASSCF wavefunction were greater than 0.05. The multireference analogue of Davidson's correction⁴⁰ was computed to estimate the importance of higher-order excitations. This correction to the CCI energy is given by

\[ Q = \Delta E (1 - C_0^2) / C_0^2 \]  

where \( \Delta E \) is the CI energy minus the reference energy, and \( C_0^2 \) is the sum of the squares of the coefficients of the reference configurations in the CI wavefunction.

The calculations using the smaller basis set were carried out on the NASA Ames Cray X-MP/48. These calculations used the MOLECULE – SWEDEN⁴¹,⁴² system of programs. The calculations using the larger basis set were carried out at Sandia National Laboratory on a Cray X-MP/416.

### III. Ab Initio Results

Table II. lists the parameters of the N₂ molecular potential energy surface which were computed with the [4s3p2d1f] and the [5s4p3d2f] basis sets. In order to be consistent with the HN₂ analysis, the diatomic calculations included the N 2p derived orbitals and the N₂ 2σ_u orbital in the active space. The CCI reference list was larger than in the HN₂ case, consisting of all the configurations generated from the N 2p derived levels. As the Table indicates, the smaller basis set yields values for
$R_e$, $\omega_e$, and $D_0$ which differ from the corresponding experimental quantities by 0.022 a$_0$, 73 cm$^{-1}$, and 13 kcal mol$^{-1}$ (based on a Dunham analysis of a quadratic fit near $R_e$). The value of $D_0$ computed in this work differs from that previously calculated with a segmented basis set of the same size by 0.5 kcal mol$^{-1}$. This earlier study concluded that the error resulting from the externally contracted CI procedure is expected to be only 0.16 kcal mol$^{-1}$ (determined by comparing CCI + Q and SDCI + Q results). Consequently, limitations in the basis set itself are responsible for the remaining difference between the value of $D_0$ calculated in this work and the experimental value.

The value of $\omega_e$ for N$_2$ computed in this study is 248 cm$^{-1}$ smaller than that obtained from the MP3/6-31G** calculations of Curtiss et al., and underestimates the experimental harmonic frequency by approximately 3.1% compared with an overestimate of approximately 7.4% by the Møller-Plesset perturbation theory calculations. Curtiss and co-workers suggested that the discrepancy between their calculated N$_2$ frequency and the experimental value is a consequence of limitations in the 6-31G** basis set. This conclusion is consistent with the earlier work of Siegbahn. Additionally, as Bartlett and Purvis, and more recently, Laidig et al. have discussed, the multiply-bonded N$_2$ system cannot be represented satisfactorily using only a single-reference, many-body perturbation theory (MBPT) technique. Consequently, it appears that limitations in both the basis set and the computational methodology have contributed to the anomalous N$_2$ frequency reported by Curtiss and co-workers. It should be noted, however, that even the CASSCF/CCI calculations reported in this work, utilizing a good quality ANO basis set, fail to represent precisely the experimental values of both $D_0$ and the harmonic stretching frequency of the N$_2$ triple bond system.

The larger basis set, [5s4p3d2f], reduces these errors to ......
The first three columns of Table III summarize the computed geometries and harmonic vibrational frequencies at the stationary points of the H + N₂ potential energy surface (using the [4s3p2d1f/3s2p1d] basis set). The values of R_{NN} and $\omega_2$ at the H + N₂ minimum were derived from supermolecule calculations and are very similar to the values of R_{e} and $\omega_e$ obtained for N₂ with the [4s3p2d1f] basis set (see Table II). These results differ because the diatomic calculations used all possible reference configurations from the active space to estimate electron correlation, while the supermolecule calculations utilized a set of selected reference configurations from the same active space. Energetically, the HN₂ species lies 3.01 kcal mol⁻¹ above the H + N₂ asymptote, while the saddle point for hydrogen atom addition is 15.16 kcal mol⁻¹ above H + N₂. Thus, there is a barrier of 12.15 kcal mol⁻¹ to the dissociation of HN₂ to H + N₂. The addition of zero-point energy corrections reduces this barrier to 8.49 kcal mol⁻¹. Single-point calculations with the [5s4p3d2f/4s3p2d] basis set (at the stationary point geometries determined with the [4s3p2d1f/3s2p1d] basis set) indicate that the HN₂ species and the saddle point are 3.73 kcal mol⁻¹ and 15.01 kcal mol⁻¹, respectively, above the H + N₂ asymptote, thus yielding a dissociation barrier for HN₂ of 11.28 kcal mol⁻¹.

The harmonic vibrational frequencies of Table III are based on normal mode analyses performed at the stationary points of the H + N₂ potential energy surface. $\omega_1$, $\omega_2$, and $\omega_3$ are the frequencies of the normal modes which correspond principally to the NH stretching motion, the NN stretching motion, and the HNN bending motion. The value of $\omega_2$ decreases from 2303 cm⁻¹ at the H + N₂ asymptote, to 2072 cm⁻¹ at the H—N₂ saddle point, and finally, to 1583 cm⁻¹ at the HN₂ minimum. Simultaneously, the value of $\omega_1$ (the frequency which essentially corresponds to the reaction coordinate for hydrogen atom addition to N₂), changes from an imaginary value (1662i cm⁻¹) at the H—N₂ saddle point to a value of 2744
cm$^{-1}$ at the HN$_2$ minimum. Finally, the value of $\omega_3$ increases from 771 cm$^{-1}$ at the H–N$_2$ saddle point to 1070 cm$^{-1}$ at the HN$_2$ minimum. These changes are consistent with the breaking of on NN $\pi$ bond and simultaneous formation of an NH $\sigma$ bond as the HN$_2$ species forms.

Table III. also compares the results obtained in the present study at the stationary points of the H + N$_2$ surface with those reported by Curtiss et al.$^{35}$ The current calculations suggest that the HN$_2$ species is energetically more stable than predicted by the earlier work, lying only 3.01 kcal mol$^{-1}$ rather than previously calculated 14.4 kcal mol$^{-1}$ above the H + N$_2$ asymptote. In addition, the barrier to dissociation (before zero-point energy corrections are added) is currently found to be 12.15 kcal mol$^{-1}$ compared with the earlier value of 10.6 kcal mol$^{-1}$. Much of this difference can be attributed to the larger basis set employed in the CASSCF/CCI study, since the basis set used by Curtiss and co-workers did not contain f functions on nitrogen or d functions on hydrogen. The d and f functions, which describe higher angular momentum states, are essential to the calculation of an accurate $D_e$ value for the HN$_2$ species.

As Table III. indicates, there are significant differences among the normal mode frequencies calculated in the current work and in that of Curtiss and co-workers. The values of $\omega_3$ (primarily the HNN bending motion) computed in the two studies are in good agreement at both the HN$_2$ minimum and the saddle point of the H + N$_2$ potential energy surface. However, this is not true of the other two normal mode frequencies. The calculated values of $\omega_1$ (primarily the NH stretching motion) and $\omega_2$ (primarily the NN stretching motion) differ substantially at the stationary points. Further, while the study of Curtiss and co-workers reports a puzzling increase in the $\omega_2$ frequency at both the saddle point and at the HN$_2$ minimum compared to its value in the N$_2$ molecule despite a lengthening of the NN bond, the $\omega_2$ frequency
does not exhibit such anomalous behavior in the CASSCF/CCI calculations. In the current study, as the NN bond lengthens, the $\omega_2$ frequency decreases monotonically. The unexplained behavior of the $\omega_2$ frequency in the work of Curtiss et al. suggests a dramatic breakdown of the Möller-Plesset perturbation expansion. As noted above, this effect is not surprising since the $N_2$ triple bond requires a multireference description to correctly represent the elongation of the NN bond.

Finally, we determined the minimum energy path (MEP) for the addition of H atom to $N_2$ by computing a series of two-dimensional grids defined in terms of $R_{NN}$ and $\theta$ (the HNN angle) for various values of $R_{NH}$ ($R_{NH} = 4.5, 4.0, 3.5, 3.0, 2.8, 2.6, 2.4, 2.2, 2.0, \text{and } 1.8 \ a_0$). A quadratic polynomial in $R_{NN}$ and $\theta$ was then fitted to each of these calculated grids (using a linear least-squares procedure), and the minima of the resulting polynomials located. While all the computed energy values are contained in the appendix, Table IV. lists the coordinates, the corresponding potential energy values, and the normal mode frequencies orthogonal to the MEP (computed from the quadratic polynomials). Figure 1 depicts the energy along the MEP as a function of the $R_{NH}$ distance. At large $R_{NH}$ distances, the H atom initially approaches $N_2$ at an angle of $\approx 119^\circ$, with $\theta$ decreasing to $\approx 115^\circ$ as $R_{NH}$ assumes its equilibrium value in HN$_2$. Simultaneously, the NN bond gradually elongates as the $R_{NH}$ distance decreases. These features of the MEP are similar to those observed in the addition of H atom to O$_2$.$^{47}$

IV. Dynamical Results

From the point of view of reaction dynamics, the lifetime of a species is perhaps the most elementary molecular property open to investigation. In the case of an extremely reactive intermediate, this parameter plays a crucial role in determining the ability of the species to participate in subsequent chemical reactions. We now
present a relatively straightforward, two-step analysis of the HN2 lifetime, utilizing both Conventional Transition State Theory in its thermodynamic formulation, and an Eckart barrier formalism which estimates the role of one-dimensional quantum mechanical tunneling in the unimolecular decay process. This analysis uses information from the stationary points and the MEP of the molecular potential energy surface, but does not attempt to describe the microscopic details of the HN2 decomposition process. An investigation of these details requires a more comprehensive representation of the potential energy surface and is left to future work. Rather, our intent is to obtain a new estimate of the HN2 lifetime based on the current CASSCF/CCI calculations, to compare this estimate with a previous determination of the HN2 lifetime, and to assess the implications of these lifetime estimates for the chemistry of the HN2 species.

The upper portion of Table V. contains a summary of the thermodynamic properties of the H + N₂ → HN₂ reaction based on information determined at the stationary points of the potential energy surface and calculated with statistical mechanics. This data is divided into two sections, the first reflects the changes in the thermodynamic state functions between the HN₂ minimum and the H + N₂ asymptote, while the second tabulates the same changes between the HN₂ minimum and the transition state (quantities marked by *). In both cases the HN₂ minimum has been chosen as the zero of energy. The data is from three distinct sources: (1) column one is based on the current CASSCF/CCI calculations using the [4s3p2df] basis set; (2) column two contains results from recent Bond Additivity Correction with fourth-order Møller-Plesset perturbation theory (BAC-MP4) calculations of Melius and Binkley; (3) the remaining columns summarize the results of the MP3/MP4 calculations of Curtiss and co-workers.

Before discussing the HN₂ lifetimes computed from the thermodynamic data us-
ing Conventional Transition State Theory, several characteristics of the data itself
deserve comment. While columns three and four list the results reported by Curtiss
et al. (Table 6. of reference 35.), columns five and six contain the same quantities
calculated directly from the energetics, geometries, and frequencies of reference 35.
In the first section of the data (relating changes between the HN₂ minimum and
the H + N₂ asymptote), there are minor discrepancies (several tenths of a kcal
mol⁻¹) between the reported and the calculated values of ΔH at both 0 K and
298.15 K. These differences arise from the apparent use by Curtiss and co-workers
of both an N₂ frequency and bond length near the experimental values (2359 cm⁻¹
and 2.075 a₀) to prepare their reported data, compared with the MP3/6-31G**
harmonic frequency of 2534 cm⁻¹ and computed bond length of 2.090 a₀ used
to generate the corresponding calculated MP3/MP4 values. We have utilized the
MP3/6-31G** harmonic frequency and bond length in preparing the calculated re-
results to allow a more direct comparison with the thermodynamic quantities based
on the CASSCF/CCI calculations. There is, however, a more serious discrepancy
between the reported and calculated values of ΔS; they differ in both magnitude
and sign. While two different computational approaches have yielded consistent
MP3/MP4 values of ΔS which agree with both the CASSCF/CCI and BAC-MP4
(computed independently) results, we have been unable to identify a reason for the
differences observed in Table V. Since the a value of ΔS is not used in the Transi-
tion State Theory calculations of the HN₂ lifetime, the conclusions of Curtiss and
co-workers are not affected by this apparent anomaly in the reported ΔS value.
The remaining thermodynamic data of Table V. (relating changes between the HN₂
minimum and the transition state) exhibit some variations between the reported
and calculated MP3/MP4 values which do affect the computed HN₂ lifetimes, but
only in a relatively minor way. Again, the conclusions of Curtiss et al. are not
affected by these small changes.

The lower portion of Table V. summarizes the classical (that is, no quantum mechanical tunneling) HN$_2$ lifetimes computed with conventional Transition State Theory from the expression:

$$\tau = \left( \frac{h}{kT} \right) \exp\left( -\frac{\Delta S^0_T}{R} + \frac{\Delta H^0_T}{RT} \right)$$

(10)

where $h$ is Planck's constant, $k$ is Boltzmann's constant, $R$ is the gas constant, and $T$ is temperature (equation (10) is simply the reciprocal of the Transition State Theory rate coefficient for the unimolecular decay of HN$_2$). The values reported by Curtiss and co-workers are consistent with the calculated MP4 results listed in the Table, while the lifetimes predicted by the CASSCF/CCI data are approximately two orders of magnitude greater at 298.15 K and approximately 19 orders of magnitude greater at 30.0 K. These differences are a simple consequence of the disparate barriers to unimolecular decay predicted by the two sets of ab initio data. At this approximate level of theory it is inappropriate to attach significance to these differences or to the precise numerical values, and the general conclusions of the two studies, based on conventional Transition State Theory calculations, are the same. First, the lifetime of the HN$_2$ species at room temperature is very short (on the order of tenths to thousandths of a microsecond); secondly, at very low temperature, the predicted very long lifetime of HN$_2$ may permit experimental observation of the species.

Up to this point in our analysis, the HN$_2$ species has been treated in a purely classical manner, ignoring quantum mechanical effects. This approach has allowed a direct comparison with a portion of the study by Curtiss and co-workers. However, since the ab initio calculations discussed earlier indicate that the HN$_2$ minimum
lies at least 3 kcal mol\(^{-1}\) above the H + N\(_2\) asymptote, every rovibrational state supported by the HN\(_2\) well is only quasi-bound. Consequently, quantum mechanical tunneling plays an essential role in the unimolecular decay of HN\(_2\). Curtiss and co-workers accounted for these quantal effects by computing the tunneling frequency of a bound particle through a parabolic barrier. In this study we have used an Eckart function to model the vibrationally adiabatic barrier to HN\(_2\) dissociation, computing HN\(_2\) lifetimes based on an estimate of the one-dimensional quantal tunneling through that barrier. The vibrationally adiabatic potential was determined by adding the zero-point energy contribution of the NN stretching motion and the HNN bending motion (frequencies listed in Table IV.) to the computed potential energy along the MEP. Further, because all of the rovibrational states of the HN\(_2\) species appear to be only quasi-bound, it does not seem appropriate to compute unimolecular lifetimes by assuming a thermally-averaged Boltzmann distribution (as does the thermodynamic formulation of Transition State Theory) created by multiple collisions of the HN\(_2\) species with other molecules in the gas-phase. Instead, the formalism of W. H. Miller\(^{39}\) was utilized to calculate unimolecular lifetimes as a function of the energy of each HN\(_2\) quasi-bound vibrational state (microcanonical lifetimes). For the purposes of this study, we have assumed that J = 0 (that is, the HN\(_2\) species is rotationally cold), although the formalism is not limited to this case alone.

Table VI. summarizes the calculated unimolecular lifetimes of HN\(_2\) as a function of the energy of each quasi-bound vibrational state. In order to provide a careful estimate of the quantal tunneling effects, the calculations reported in Table VI. represent the results of three distinct Eckart Barrier models (designated Eckart Barrier-1, Eckart Barrier-2, and Eckart Barrier-3). In these models, the three parameters which completely determine the Eckart function, the height of the barrier
with respect to HN$_2$ ($V_0$), the height of the barrier with respect to H + N$_2$ ($V_1$), and the imaginary frequency (which determines the curvature at the top of the barrier) ($\omega^*$), were systematically varied in an attempt to reflect accurately the properties of the HN$_2$ potential energy surface.

In the first model, Eckart Barrier-1, no effort was made to fit numerically the Eckart function to the shape of the vibrationally adiabatic potential. Rather, the parameters $V_0$, $V_1$, and $\omega^*$ were set equal to the zero-point-corrected (adding the zero-point energy of the NN stretch and the HNN bend) ab initio potential energy values and to the calculated imaginary frequency. Table VI. reports three calculations of the HN$_2$ lifetime as a function of energy using the Eckart Barrier-1 model: (1) a calculation using the ab initio results of the [4s3p2d1f] basis set, with $V_0 = 12.417$ kcal mol$^{-1}$, $V_1 = 15.931$ kcal mol$^{-1}$, and $\omega^* = 1662.2i$ cm$^{-1}$; (2) a calculation using the ab initio results of the [5s4p3d2f] basis set (at the [4s3p2d1f] geometry), with $V_0 = 11.549$ kcal mol$^{-1}$, $V_1 = 15.777$ kcal mol$^{-1}$, and $\omega^* = 1662.2i$ cm$^{-1}$; (3) a calculation using the MP4/6-31G** results of Curtiss and co-workers with $V_0 = 10.310$ kcal mol$^{-1}$, $V_1 = 26.350$ kcal mol$^{-1}$, and $\omega^* = 1533i$ cm$^{-1}$. Figure 2. compares the Eckart Barrier-1 model to the vibrationally adiabatic potential computed with the [4s3p2d1f] basis set. As Table VI. indicates, the MP4/6-31G** potential energy surface of Curtiss and co-workers supports only two quasi-bound vibrational states, while the potential energy surfaces computed in this work with the [4s3p2d1f] basis set and the [5s4p3d2f] basis set support six and five quasi-bound vibrational states, respectively. The Table also demonstrates the dynamical effect of using a larger ab initio basis set on the HN$_2$ lifetimes which range from $6.37 \times 10^{-9}$ seconds to $1.76 \times 10^{-13}$ seconds for the six vibrational levels of the [4s3p2d1f] basis, and from $1.25 \times 10^{-9}$ seconds to $1.02 \times 10^{-13}$ seconds for the five vibrational levels of the [5s4p3d2f] basis. Finally, the lifetime computed for HN$_2$
dissociation from the lowest vibrational state of the [4s3p2d1f] potential energy surface is approximately two orders of magnitude greater than the lifetime calculated for the unimolecular decay of HN$_2$ from the lowest vibrational state supported by the MP4/6-31G** potential energy surface.

However, as Figure 2. indicates, the Eckart Barrier-1 model both significantly overestimates the width of the barrier and does not closely approximate the shape of the vibrationally adiabatic potential. We therefore attempted to define a single model which is able to represent accurately both the width of the barrier and the shape of the vibrationally adiabatic potential. This proved to be impossible. Consequently, two additional models were constructed by numerically fitting (in a non-linear, least-squares sense) an Eckart function to the vibrationally adiabatic potential determined with the [4s3p2d1f] basis set. In the first of these new models, designated Eckart Barrier-2, $V_0$ was held fixed at the value used in the Eckart Barrier-1 model, while $V_1$ and $\omega^*$ were treated as adjustable parameters. The fitted Eckart Barrier-2 model is characterized by the parameters $V_0 = 12.417$ kcal mol$^{-1}$, $V_1 = 12.997$ kcal mol$^{-1}$, and $\omega^* = 2441.3i$ cm$^{-1}$, but, as Figure 3. indicates, it significantly underestimates the width of the barrier. This model does, however, possess the attribute of closely approximating the shape of the HN$_2$ side of the vibrationally adiabatic potential. The second of the new models, identified as Eckart Barrier-3, was constructed by holding both $V_0$ and $V_1$ fixed at the values which characterize the Eckart Barrier-2 model (hence, $V_0$ is the same value in all three models considered in this work) and treating $\omega^*$ as the only adjustable parameter in the non-linear least-squares fitting procedure. The resulting Eckart Barrier-3 model is defined by the parameters $V_0 = 12.417$ kcal mol$^{-1}$, $V_1 = 12.997$ kcal mol$^{-1}$, and $\omega^* = 1715.6i$ cm$^{-1}$, and is compared in Figure 4. to the vibrationally adiabatic potential. While the Eckart Barrier-3 model overestimates the width of
the barrier, it does succeed in closely approximating the shape of $H + N_2$ side of
the vibrationally adiabatic potential.

The two models provide a lower bound (Eckart Barrier-2) and an upper bound
(Eckart Barrier-3) for the $HN_2$ unimolecular lifetime of each quasi-bound vibrational
state associated with the [4s3p2d1f] basis set. These bounds are summarized in the
final two columns of Table VI. Note that as the top of the barrier is approached,
the widths of the two model barriers become increasingly similar. Consequently,
the difference between the two lifetimes computed for each quasi-bound vibrational
state decreases with increasing energy. In the case of the lowest-lying quasi-bound
vibrational state, the two computed lifetimes differ by approximately a factor of 66;
for the highest-lying state, they differ only by approximately a factor of 1.4. The
$HN_2$ lifetimes predicted by these two models, which include the effects of quantum
mechanical tunneling, suggest that the $HN_2$ species is extremely short-lived, even
at very low temperatures. As a result, we expect that it will be difficult to observe
the $HN_2$ species experimentally, and that the role of $HN_2$ in both termolecular
recombination processes and combustion processes will be dramatically limited by
its very short lifetime.

V. Conclusions

This study has focused on two aspects of the addition of atomic hydrogen to $N_2$:
first, a characterization of the minimum energy path for the addition reaction, and
secondly, a determination of the unimolecular lifetime of the $HN_2$ species. While
this work has presented neither a global description of the potential energy surface
for the addition reaction nor a discussion of the microscopic details of the dynamics,
it has reached several important conclusions:

(1) The $HN_2$ minimum is found to lie approximately 3.01 kcal mol$^{-1}$ above the
H + N₂ asymptote, compared to 14.4 kcal mol⁻¹ calculated earlier by Curtiss and co-workers.

(2) While this earlier study computed a zero-point-corrected barrier to dissociation of approximately 5.8 kcal mol⁻¹, the current work has estimated this barrier to be approximately 8.5 kcal mol⁻¹.

(3) The current calculations indicate that six quasi-bound vibrational levels exist in the HN₂ well, compared to the two quasi-bound levels predicted in the earlier study of Curtiss and co-workers.

(4) The current work has used both a larger basis set and an improved methodology to calculate a new estimate of the NN harmonic stretching frequency in the N₂ and HN₂ species.

(5) We estimate that the lifetime of the lowest-lying quasi-bound vibrational state of HN₂ to lie between 8.8 × 10⁻¹¹ seconds and 5.8 × 10⁻⁹ seconds. This predicted lifetime is between a factor of two and a factor of one hundred longer than the lifetime of the HN₂ species predicted by Curtiss and co-workers.

(6) However, we conclude, in concert with the conclusions of Curtiss and co-workers, that it will be very difficult to observe the HN₂ species experimentally, and that the predicted short lifetime of the HN₂ species will severely circumscribe its role in both termolecular reaction mechanisms and proposed combustion mechanisms.

ACKNOWLEDGEMENTS

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42. SWEDEN is a vectorized SCF-MCSCF-direct CI, conventional CI-CPF-MCPF program written by P. E. M. Siegbahn, C. W. Bauschlicher, Jr., B. Roos, P. R. Taylor, A. Heiberg, J. Almlöf, S. R. Langhoff, and D. P. Chong.


Table I. Qualitative Character of the CASSCF Orbitals for $H + N_2 \rightarrow HN_2^*$

<table>
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<tr>
<th>orbital</th>
<th>$H + N_2$</th>
<th>$HN_2$</th>
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<td>1a'</td>
<td>N1s</td>
<td>N1s</td>
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<td>N1s</td>
</tr>
<tr>
<td>3a'</td>
<td>N2s</td>
<td>N2s</td>
</tr>
<tr>
<td>4a'</td>
<td>N2s</td>
<td>N2s</td>
</tr>
<tr>
<td>5a'</td>
<td>N$_2$3σ$_g$</td>
<td>N$_2$3σ$_g$</td>
</tr>
<tr>
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<td>N$_2$1π$_u$</td>
<td>N$_2$1π$_u$ + H1s</td>
</tr>
<tr>
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<td>H1s</td>
<td>N$_2$1π$_g$ + H1s</td>
</tr>
<tr>
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<td>N$_2$1π$_g$</td>
<td>N$_2$1π$_g$ - H1s</td>
</tr>
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<td>N$_2$3σ$_u$</td>
<td>N$_2$3σ$_u$</td>
</tr>
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<td>N$_2$1π$_u$</td>
<td>N$_2$1π$_u$</td>
</tr>
<tr>
<td>2a''</td>
<td>N$_2$1π$_g$</td>
<td>N$_2$1π$_g$</td>
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Table II. Computed Potential Curve Parameters for N₂.

<table>
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<tr>
<th></th>
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<td>( R_e )</td>
<td>2.096</td>
<td></td>
<td>2.089</td>
<td>2.074</td>
</tr>
<tr>
<td>( \omega_e )</td>
<td>2286</td>
<td></td>
<td>2269</td>
<td>2359</td>
</tr>
<tr>
<td>( D_0 )</td>
<td>212</td>
<td>217</td>
<td>219</td>
<td>225.0</td>
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</table>

\( R_e \) in \( a_0 \), \( \omega_e \) in \( \text{cm}^{-1} \), \( D_0 \) in \( \text{kcal mol}^{-1} \)
Table III. Stationary Points on the II + N₂ Surface.\textsuperscript{a}

<table>
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<th></th>
<th>this work</th>
<th>CDP\textsuperscript{b}</th>
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<tr>
<td></td>
<td>II + N₂</td>
<td>II—N₂</td>
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<tr>
<td></td>
<td>minimum</td>
<td>saddle point</td>
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<tr>
<td>( R_{NH} )</td>
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<td>( R_{NN} )</td>
<td>2.095</td>
<td>2.146</td>
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<tr>
<td>( \theta )</td>
<td>118.6</td>
<td>116.3</td>
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<tr>
<td>( \omega_1 )</td>
<td>1662\textit{i}</td>
<td>2744</td>
</tr>
<tr>
<td>( \omega_2 )</td>
<td>2303</td>
<td>2072</td>
</tr>
<tr>
<td>( \omega_3 )</td>
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<td>( \Delta E^c )</td>
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<td>( \Delta E_{corr} )</td>
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\textsuperscript{a} Energies are in kcal mol\textsuperscript{-1}, bond lengths in \( a_0 \), angles in degrees, and vibrational frequencies are in cm\textsuperscript{-1}. This work used the [4s3p2d1f/3s2p1d] basis set.

\textsuperscript{b} Curtiss, Drapcho, and Pople, reference 35.

\textsuperscript{c} Energy relative to HN₂.

\textsuperscript{d} Based on MP4/6-31G** energies evaluated at MP3/6-31G** geometries.

\textsuperscript{e} Zero-Point Energy
Table IV. Computed Minimum Energy Path for H + N₂.\(^a\)

<table>
<thead>
<tr>
<th>R(_{NH})</th>
<th>R(_{NN})</th>
<th>(\theta)</th>
<th>(\omega_2)</th>
<th>(\omega_3)</th>
<th>E(CCl + Q) (hartree)</th>
<th>(\Delta E^b) (kcal mol(^{-1}))</th>
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<td>1202</td>
<td>1659</td>
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<td>3.25</td>
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\(^a\) Bond lengths are in a₀, angles are in degrees, and frequencies in cm\(^{-1}\).

\(^b\) Energy relative to HN₂ minimum.
Table V. Thermodynamic Properties and Unimolecular Lifetimes for HN$_2$ → H + N$_2$

<table>
<thead>
<tr>
<th>Thermodynamic Data</th>
<th>CASSCF</th>
<th>BAC-MP$^a$</th>
<th>MP3 (reported$^b$)</th>
<th>MP4</th>
<th>MP3 (calculated$^c$)</th>
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<td>$\Delta E$</td>
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<tr>
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<tr>
<td>$\Delta H^\dagger_{30}$</td>
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<td>-</td>
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<td>$\Delta S^\dagger_{298}$</td>
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<tr>
<td>30.0 K</td>
<td>9.53 $\times$ 10$^{49}$</td>
<td>1.1 $\times$ 10$^{31}$</td>
<td>3.16 $\times$ 10$^{52}$</td>
<td>1.42 $\times$ 10$^{31}$</td>
</tr>
<tr>
<td>298.15 K</td>
<td>2.08 $\times$ 10$^{-7}$</td>
<td>2.2 $\times$ 10$^{-9}$</td>
<td>3.69 $\times$ 10$^{-7}$</td>
<td>2.62 $\times$ 10$^{-9}$</td>
</tr>
</tbody>
</table>

$^a$ Melius and Binkley, reference 48.
$^b$ Curtiss, Drapcho, and Pople, reference 35.
$^c$ See text.
$^\dagger$ Refers to thermodynamic quantities of the transition state; remaining thermodynamic quantities are those of the H + N$_2$ asymptote. Both sets of thermodynamic data are calculated with respect to the HN$_2$ minimum. Energies and enthalpies are in kcal mol$^{-1}$, entropies in cal mol$^{-1}$ deg$^{-1}$, and lifetimes in seconds.
Table VI. Calculated Unimolecular Lifetimes\(^a\) of II\(_2\)

<table>
<thead>
<tr>
<th>Quasi-Bound State(^b)</th>
<th>Eckart Barrier-1</th>
<th>Eckart Barrier-2</th>
<th>Eckart Barrier-3</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>[4s3p2d1f]</td>
<td>[5s4p3d2f]</td>
<td>[4s3p2d1f]</td>
</tr>
<tr>
<td>7.72</td>
<td>(6.37 \times 10^{-9})</td>
<td>(1.25 \times 10^{-9})</td>
<td>(8.82 \times 10^{-11})</td>
</tr>
<tr>
<td>10.78</td>
<td>(5.95 \times 10^{-11})</td>
<td>(1.46 \times 10^{-11})</td>
<td>(4.30 \times 10^{-12})</td>
</tr>
<tr>
<td>12.24</td>
<td>(7.95 \times 10^{-12})</td>
<td>(2.14 \times 10^{-12})</td>
<td>(1.19 \times 10^{-12})</td>
</tr>
<tr>
<td>13.84</td>
<td>(1.04 \times 10^{-12})</td>
<td>(3.30 \times 10^{-13})</td>
<td>(3.42 \times 10^{-13})</td>
</tr>
<tr>
<td>15.30</td>
<td>(2.18 \times 10^{-13})</td>
<td>(1.02 \times 10^{-13})</td>
<td>(1.40 \times 10^{-13})</td>
</tr>
<tr>
<td>15.56</td>
<td>(1.76 \times 10^{-13})</td>
<td>–</td>
<td>(1.24 \times 10^{-13})</td>
</tr>
</tbody>
</table>

<table>
<thead>
<tr>
<th>Quasi-Bound State(^b)</th>
<th>Eckart Barrier-1</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>[MP4/6-31G(^{**})]</td>
</tr>
<tr>
<td>9.69</td>
<td>(7.13 \times 10^{-11})</td>
</tr>
<tr>
<td>12.77</td>
<td>(8.60 \times 10^{-13})</td>
</tr>
</tbody>
</table>

\(^a\) Lifetimes in seconds.

\(^b\) Energies of quasi-bound states in kcal mol\(^{-1}\).

\(^c\) Curtiss, Drapcho, and Pople, reference 35.
FIGURE CAPTIONS

Figure 1.: Interaction energy for the addition of a hydrogen atom to N₂ as a function of $R_{NH}$ along a minimum energy path (MEP) obtained by optimizing $R_{NN}$ and $\theta$ at each $R_{NH}$ value. The energies are from CASSCF/CCI + Q calculations (see text).

Figure 2.: Eckart Barrier-1 model. The bold solid curve is the vibrationally adiabatic barrier calculated from the CASSCF/CCI + Q data along the minimum energy path (MEP). The solid curve is an Eckart function. The horizontal dashed curves represent the quasi-bound vibrational states of the HN₂ well. In the separable harmonic approximation, the energy of each vibrational state is given by $E(n_1,n_2,n_3) = \sum_{i=1}^{3} (n_i + \frac{1}{2})\hbar\nu_i$. Each state is identified by a set of quantum numbers $(n_1, n_2, n_3)$, where $n_1 = HN \text{ stretch}$ quantum number, $n_2 = NN \text{ stretch}$ quantum number, $n_3 = HNN \text{ bend}$ quantum number.

Figure 3.: Eckart Barrier-2 model. The bold solid curve is the vibrationally adiabatic barrier calculated from the CASSCF/CCI + Q data along the minimum energy path (MEP). The solid curve is an Eckart function. The horizontal dashed curves represent the quasi-bound vibrational states of the HN₂ well. In the separable harmonic approximation, the energy of each vibrational state is given by $E(n_1,n_2,n_3) = \sum_{i=1}^{3} (n_i + \frac{1}{2})\hbar\nu_i$. Each state is identified by a set of quantum numbers $(n_1, n_2, n_3)$, where $n_1 = HN \text{ stretch}$ quantum number, $n_2 = NN \text{ stretch}$ quantum number, $n_3 = HNN \text{ bend}$ quantum number.

Figure 4.: Eckart Barrier-3 model. The bold solid curve is the vibrationally adiabatic barrier calculated from the CASSCF/CCI + Q data along the minimum energy path (MEP). The solid curve is an Eckart function. The horizontal dashed curves represent the quasi-bound vibrational states of the HN₂ well. In the separable harmonic approximation, the energy of each vibrational state is given by $E(n_1,n_2,n_3) = \sum_{i=1}^{3} (n_i + \frac{1}{2})\hbar\nu_i$. Each state is identified by a set of quantum numbers $(n_1, n_2, n_3)$, where $n_1 = HN \text{ stretch}$ quantum number, $n_2 = NN \text{ stretch}$ quantum number, $n_3 = HNN \text{ bend}$ quantum number.