LABORATORY STUDIES OF REFRACTORY METAL OXIDE SMOKES

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Studies of the properties of refractory metal oxide
smokes condensed from a gas containing various
combinations of SiH₄, Fe(CO)₅, Al(CH₃)₃, TiCl₄, O₂ and
N₂O in a hydrogen carrier stream at 500 K > T > 1500 K
have been performed in our laboratory. Ultraviolet,
visible and infrared spectra of pure, amorphous SiOₓ,
FeOₓ, AlOₓ and TiOₓ smokes will be discussed, as will
the spectra of various co-condensed amorphous oxides,
such as Fe₃SiOₓ or Fe₃AlOₓ. Preliminary studies of the
changes induced in the infrared spectra of iron-
containing oxide smokes by vacuum thermal annealing
suggest that such materials become increasingly opaque
in the near infrared with increased processing;
hydration may have the opposite effect (see Fig. 1).
More work on the processing of these materials is
required to confirm such a trend; this work is
currently in progress. Preliminary studies of the
ultraviolet spectra of amorphous SiOₓ and MgSiOₓ
smokes revealed no interesting features in the region
from 200-300 nm. Studies of the ultraviolet spectra of
both amorphous, hydrated and annealed SiOₓ, TiOₓ, AlOₓ
and FeOₓ smokes are currently in progress.

Finally, data on the oxygen isotopic composition of
the smokes produced in our experiments will be
presented which indicate that the oxygen becomes
isotopically fractionated during grain condensation.
Oxygen in our grains is as much as 3% per amu lighter
than the oxygen in the original gas stream. We are
currently conducting experiments to understand the
mechanism by which fractionation occurs. Such processes may have observable consequences in circumstellar outflows if the effects become larger as the total pressure decreases. This might be expected if the fractionation mechanism operates during the formation of the initial molecular oxides rather than during the nucleation of the refractory grains which form from such oxides.

Figure 1. Infrared spectra of amorphous Fe-Al-silicate smokes. The lower spectrum is the initial condensate. The middle spectrum is that of a sample which was immersed in liquid water for 5 days at 378 K and then freeze-dried. The top spectrum is that of a sample which has been vacuum annealed for 16 hrs. at 1200 K. Note that a scale change occurs at 2000 cm$^{-1}$. 

![Infrared spectra diagram with composition percentages: 1.3% Al2O3, 95.3% SiO2, 3.4% FeO.](image)