Program 7  Hydrogen Interactions in Aluminum-Lithium Alloys

S.W. Smith and J.R. Scully

Objective

The objective of this work is to develop a fundamental understanding of the effects of dissolved and trapped hydrogen on the mechanical properties of selected Al-Li-Cu-X alloys. We propose to: (a) distinguish hydrogen induced EAC from aqueous dissolution controlled processes, (b) correlate hydrogen induced EAC with mobile and trapped hydrogen concentrations and (c) identify significant trap sites and hydrides (if any) through the utilization of model alloys and phases.
Hydrogen Interactions in Aluminum Lithium Alloys

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Experimental Problems associated with Investigation of Hydrogen Effects in Al-Li-Cu-X Alloys

1. **Problem:** Low fracture toughness is observed for aged Al-Li-Cu-X alloys stressed in the S-T orientation. Intergranular and inter-subgranular fracture occurs in air or in vacuum making it difficult to detect hydrogen assisted fracture on the basis of fractography. **Solution:** Test specimens in L-T orientation after appropriate aging to produce shear band cracking in air or vacuum.

2. **Problem:** Low hydrogen diffusivity and solubility in aluminum alloys is compounded by an oxide permeation barrier. These factors make hydrogen analysis difficult. **Solution:** Use thermal desorption spectroscopy and Pd coated samples.

3. **Problem:** Aqueous EAC response may be dominated by dissolution assisted processes even during cathodic charging. This tends to mask hydrogen effects. Fractographic features distorted by dissolution. **Solution:** Use modification of method originally described by Gruhl and co-workers or use Pd coated breaking load samples.
Approach

1. Age alloys to produce shear band fracture in air.

2. Perform modified breaking load and slow strain rate tests using pre-charged Al-Li-Cu-X alloys in L-T orientation. Use Pd coated samples with native oxide removed by sputter etching. Use alloy with a known hydrogen response (i.e., 7075-T6) as a control.

3. Analyze fracture surfaces using "advanced" methods as well as Al-Li-Cu-X alloys specimens.

4. Conduct hydrogen analysis on hydrogenated model alloys as follows: Dehvanathan-Stuchel permeation method.

   a) Modified Devanathan-Stuchel permeation method
   b) Thermal desorption spectroscopy
   c) Hydride detection methods
   d) Nuclear methods
Fabrication of Flat Tensile Bar in L-T Orientation for Breaking Load Studies
Constant Deflection Apparatus for use in Breaking Load studies
Advantages to Using Palladium Coatings

- Can remove $\text{Al}_2\text{O}_3$ layer, which impedes hydrogen diffusion.

- Surface of specimen will not be affected by cathodic charging.

- Can distinguish between aqueous and hydrogen effects.
Etching of an Aluminum Alloy Substrate by dc Sputter Etching
Deposition of Palladium by dc Sputtering
Tensile Specimen for Modified Breaking Load Studies

Notched Surface:
Uncoated.
Exposed to lab air and/or vacuum.

Back Surface:
Coated with Pd.
Cathodically charged.
<table>
<thead>
<tr>
<th>Material</th>
<th>Orientation</th>
<th>Treatment</th>
<th>Fractography at Applied Load</th>
</tr>
</thead>
<tbody>
<tr>
<td>7075 - T6</td>
<td>L - T</td>
<td>Glove Box / Lab Air</td>
<td>slip band slip band slip band</td>
</tr>
<tr>
<td>2090 -</td>
<td>L - T</td>
<td>Glove Box / Lab Air</td>
<td>slip band slip band slip band</td>
</tr>
<tr>
<td>8090 -</td>
<td>L - T</td>
<td>Glove Box / Lab Air</td>
<td>slip band slip band slip band</td>
</tr>
<tr>
<td>7075 - T6</td>
<td>L - T</td>
<td>Cathodic Charging</td>
<td>to be determined</td>
</tr>
<tr>
<td>2090 -</td>
<td>L - T</td>
<td>Cathodic Charging</td>
<td>to be determined</td>
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<td>8090 -</td>
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<td>Cathodic Charging</td>
<td>to be determined</td>
</tr>
</tbody>
</table>
## Hydrogen Analysis Methods Pertinent to Aluminum-Lithium Alloys

<table>
<thead>
<tr>
<th>Method</th>
<th>Advantage</th>
<th>Facility</th>
<th>Comment</th>
</tr>
</thead>
<tbody>
<tr>
<td>Devanathan Stuchurski Permeation</td>
<td>Inexpensive Significant Experience</td>
<td>U.Va-CESE</td>
<td>Favors alloys with permeability and mobile H</td>
</tr>
<tr>
<td>Nuclear Reaction $^3$He(d,p)$^4$He</td>
<td>Absolute conc. depth profile used for Al</td>
<td>Sandia</td>
<td>Not readily avail. trapped+mobile D must use $D_2O$</td>
</tr>
<tr>
<td>Thermal Desorption Spectroscopy</td>
<td>relative conc. Spectroscopic assess trap strength</td>
<td>U.Va-CESE</td>
<td>proven for Al-Li must use $D_2O$</td>
</tr>
<tr>
<td>Neutron Act. Neutron Rad.</td>
<td>thick samples &quot;&quot;</td>
<td>U. Va.-Nuclear</td>
<td>quantitative qualitative</td>
</tr>
</tbody>
</table>
Hydrogen Concentration Profile During Hydrogen Permeation Studies

Charging side:

hydrogen reduction

\[ C_H = f(n_c) \]

\[ H^+ + e^- \rightarrow H \]

Exit side:

hydrogen oxidation

\[ C_H \rightarrow 0 \]

\[ H \rightarrow H^+ + e^- \]
Schematic of Devanathan-Stachurski Permeation Cell
Schematic of Differential Permeation Cell

charging side

control cell

exit side

potentiostat/galvanostat

bipotentiostat
Hydrogen Transport in Endothermic Hydrogen Absorbers (Fe) is a Strong Function of The Nature and Density of Traps

\[ K = \frac{N_i k_r}{p_r} = \exp(-E_b/RT), \quad k \text{ is the trap rate, } p \text{ is the release rate} \]

\[ N_i \text{ (lattice sites)} = 2.6 \times 10^{23} \text{ octahedral sites/cm}^3 \text{ for BCC iron} \]

\[ E_b > E_s \quad \text{Strong Trap (~29 KJ/mole)} \]
A Series of Permeation Rise and Decay Transients is Utilized to Separate Reversible from Irreversible Trapping

\[
\begin{align*}
\text{DECAY: } & \frac{(l_l - l_2)}{(l_1 - l_2)} = 1 - \frac{2}{\sqrt{\pi t}} \cdot \exp \left( -\frac{1}{4\tau} \right) \\
\text{RISE: } & \frac{(l_l - l_1)}{(l_2 - l_1)} = \frac{2}{\sqrt{\pi t}} \cdot \exp \left( -\frac{1}{4\tau} \right)
\end{align*}
\]

REARRANGING:

\[
\begin{align*}
\text{DECAY: } & \log \left( \frac{l_l - l}{\sqrt{t}} \right) = \log \frac{2F \sqrt{D} (C_1 - C_2)}{\sqrt{\pi}} - \frac{L^2 \log e}{4D} \times \frac{1}{t} \\
\text{RISE: } & \log \left( \frac{l_l - l_1}{\sqrt{t}} \right) = \log \frac{2F \sqrt{D} (C_2 - C_1)}{\sqrt{\pi}} - \frac{L^2 \log e}{4D} \times \frac{1}{t}
\end{align*}
\]

WHERE: \( \tau = Dt/L^2 \)
Diffusion coefficient of Hydrogen in various Al alloys

Temp. (K)

<table>
<thead>
<tr>
<th>823</th>
<th>773</th>
<th>723</th>
<th>673</th>
<th>623</th>
<th>573</th>
<th>523</th>
<th>473</th>
<th>423</th>
<th>373</th>
<th>323</th>
<th>273</th>
</tr>
</thead>
</table>

$D$ (cm$^2$/sec)

- $\text{Al}-1\% \text{ Li}$
- $\text{Al}-2\% \text{ Li}$
- $\text{Al}-3\% \text{ Li}$

- 99.99% Al
- Al-2Li extrapolated to RT

Anyalebechi

Watson & Meshii
for 99.99% Al
Design of Thermal Desorption Spectroscopy system

Thermal Desorption Spectroscopy Offers Several Advantages

Constant Temperature

Temperature Ramp
Solubility of Hydrogen in Al alloys at 10,000 atm.

Temp. (K)

10^{-3}

10^{-4}

10^{-5}

10^{-6}

10^{-7}

10^{-8}

10^{-9}

10^{-10}

10^{-11}

10^{-12}

\frac{1}{T \times 10^4} \text{ K}^{-1}

Al 99.99%

Al - 1% Li

Al - 2% Li

Al - 3% Li
# Hydrogen Trapping

<table>
<thead>
<tr>
<th>Model Alloy</th>
<th>Exploitable Trap Site</th>
<th>Possible Trapping - Interaction</th>
</tr>
</thead>
<tbody>
<tr>
<td>Al - 3Li</td>
<td>$\delta$, $\delta^\prime$, Li$_{ss}$</td>
<td>$\delta^\prime$-interphasal; $\delta, \delta^\prime$-hydride; Li-solute</td>
</tr>
<tr>
<td>Al - 3Cu</td>
<td>$\theta^\prime$, $\theta^\prime$, $\theta$, Cu$_{ss}$</td>
<td>$\theta^\prime, \theta^\prime$-interphasal; $\theta$-void</td>
</tr>
<tr>
<td>Al - Li-Cu</td>
<td>$\delta^\prime$, $T_1$, $T_2$</td>
<td>$\delta^\prime, T_1, T_2$-hydride; $T_1$-interphasal; $T_2$-void</td>
</tr>
<tr>
<td>T$_1$</td>
<td>T$_1$</td>
<td>T$_1$-hydride</td>
</tr>
<tr>
<td>Al</td>
<td>g.b., vacancies, voids, microvoids</td>
<td>interfacial, point defects, voids</td>
</tr>
<tr>
<td>Al - Zr</td>
<td>$\beta$, $\beta^\prime$</td>
<td>$\beta$-void; $\beta^\prime$-interphasal; $\beta, \beta^\prime$-hydrides</td>
</tr>
<tr>
<td>Al$_3$Zr ($\beta^\prime$)</td>
<td>$\beta^\prime$</td>
<td>$\beta^\prime$-hydride</td>
</tr>
<tr>
<td>Al - Li-Zr</td>
<td>$\delta^\prime$ coats $\beta^\prime$</td>
<td>interphasal; $\delta^\prime, \beta^\prime$-hydride</td>
</tr>
</tbody>
</table>
1. Hydrogen permeation cells assembled. Thin foils must be prepared.

2. Breaking load configuration built and tested. Specimen exposures to begin this summer.

3. Thermal desorption system in design stages. Equipment purchases to follow.

4. Hydrogen evolution reaction kinetics studies to be undertaken in July to ascertain hydrogen production capability of model alloys and phases.