CALCULATION OF PHASE EQUILIBRIA IN THE Y₂O₃-Yb₂O₃-ZrO₂ SYSTEM

Nathan S. Jacobson  
NASA Glenn Research Center  
Cleveland, OH 44135

Zi-Kui Liu  
Department of Materials Science and Engineering  
Pennsylvania State University  
State College, PA 16802

Larry Kaufman  
Brookline, MA 02146

Fan Zhang  
Computherm LLC  
Madison, WI 53719

ABSTRACT

Rare earth oxide stabilized zirconias find a wide range of applications. An understanding of phase equilibria is essential to all applications. In this study, the available phase boundary data and thermodynamic data is collected and assessed. Calphad-type databases are developed to completely describe the Y₂O₃-ZrO₂, Yb₂O₃-ZrO₂, and Y₂O₃-Y₂O₃ systems. The oxide units are treated as components and regular and sub-regular solution models are used. The resultant calculated phase diagrams show good agreement with the experimental data. Then the binaries are combined to form the database for the Y₂O₃-Yb₂O₃-ZrO₂ pseudo-ternary.

INTRODUCTION

Yttria-stabilized zirconia is one of the most widely used ceramics, with applications ranging from thermal barrier coatings to solid oxide fuel cells. There has been interest in using other rare earth oxides alone or in combination to stabilize zirconia. These different combinations offer improved properties, such as reduced thermal conductivity.

A survey of phase equilibria data for rare-earth oxide/zirconia systems shows a good deal of information on the Y₂O₃-ZrO₂ system, but only limited information on the other rare-earth oxide/zirconia systems (1-3). Experimental data on two or more rare earth oxides in combination with zirconia is even more limited. For this reason, a Calphad (CAalculation of PHase Diagrams) assessment is valuable (4). In this method, a series of binary databases are combined to form reliable high order databases.

Due to the widespread use of the Y₂O₃-ZrO₂ binary system, there are several assessments in the literature. Kaufman (5) first approached this system by treating each oxide unit as a component. Degtyarev and Voronin (6-7) further simplified this system...
by treating the two cubic phases—cubic ZrO2 and cubic Y2O3—as one cubic phase with a miscibility gap. Du et al. (1) have thoroughly reviewed all available literature on this system through 1990 and developed a Calphad-type database with this miscibility gap simplification. The group of Yokokawa (8-9) has also calculated the Y2O3-ZrO2 phase diagram and extended these calculations to other rare-earth oxide/zirconia systems.

In this paper we re-assess the Y2O3-ZrO2 system. We do not use the miscibility gap simplification, because cubic ZrO2 and cubic Y2O3 are of different space groups. Also a miscibility gap leads to convergence problems from the corresponding multi-component systems. We also assess the Yb2O3-ZrO2 and Y2O3-Yb2O3 systems and combine these to form the Y2O3-Yb2O3-ZrO2 ternary.

**APPRAOCH**

The first step in the assessment process is to survey all available data in the literature. The phases in the Y2O3-Yb2O3-ZrO2 system are given in Table I.

<table>
<thead>
<tr>
<th>Phase</th>
<th>Abbreviation</th>
</tr>
</thead>
<tbody>
<tr>
<td>Monoclinic ZrO2</td>
<td>mz</td>
</tr>
<tr>
<td>Tetragonal ZrO2</td>
<td>tz</td>
</tr>
<tr>
<td>Cubic ZrO2</td>
<td>cz</td>
</tr>
<tr>
<td>Liquid Y2O3-Yb2O3-ZrO2</td>
<td>liq</td>
</tr>
<tr>
<td>Cubic Y2O3</td>
<td>cy</td>
</tr>
<tr>
<td>Hexagonal Y2O3</td>
<td>hy</td>
</tr>
<tr>
<td>Zr3Y4O12</td>
<td>delta y</td>
</tr>
<tr>
<td>Cubic Yb2O3</td>
<td>cyb</td>
</tr>
<tr>
<td>Hexagonal Yb2O3</td>
<td>hyb</td>
</tr>
<tr>
<td>Zr3Yb4O12</td>
<td>delta yb</td>
</tr>
</tbody>
</table>

As noted, there is a good deal of experimental information on the Y2O3-ZrO2 system. This has been discussed in detail in other publications (1-3) and will only be briefly summarized here. The liquidus line of Noguchi et al. (10) was used in this assessment. There are numerous measurements of the tetragonal/(cubic + tetragonal)/cubic phase boundaries (11-17). Several different techniques were used and, in general, these show remarkably good agreement. The studies of Stubican et al. (12) are the most complete examination of this system and were used in the assessment. Four invariant points have been reported for this system—these are listed in Table II. There is an intermediate compound of composition Zr3Y4O12, which decomposes to the cubic solid solution of ZrO2. A second, pyrochlore-type compound in the Y2O3-rich portion of the diagram has been reported by other investigators, but its existence is controversial (18) and it will not be considered here.
Table II. Invariant points and decomposition temperatures in the Y₂O₃-ZrO₂ and Yb₂O₃-ZrO₂ pseudo-binaries.

<table>
<thead>
<tr>
<th></th>
<th>Temperature (K)</th>
<th>Composition x(Y₂O₃)</th>
<th>Reference</th>
<th>Comments</th>
</tr>
</thead>
<tbody>
<tr>
<td>Eutectic liq = cz + cy</td>
<td>2688</td>
<td>0.817</td>
<td>(12)</td>
<td></td>
</tr>
<tr>
<td>Peritectic cy = liq + hy</td>
<td>2720</td>
<td>0.778</td>
<td>(12)</td>
<td>Introduced, not measured</td>
</tr>
<tr>
<td>Euctectoid cz = delta y + cy</td>
<td>1650</td>
<td>0.435</td>
<td>(13)</td>
<td></td>
</tr>
<tr>
<td>Euctectoid tz = mz + cz</td>
<td>838</td>
<td>0.039</td>
<td>(19)</td>
<td>Difficult to attain equilibria at this low temperature</td>
</tr>
<tr>
<td>Euctectoid cz = delta y + mz</td>
<td>&lt; 673</td>
<td>0.20-0.30</td>
<td>(12)</td>
<td>Difficult to attain equilibria at this low temperature Approximate euctectoid discussed in (12)</td>
</tr>
<tr>
<td>Decomposition Temperature of delta y</td>
<td>1655</td>
<td>0.4</td>
<td>(20)</td>
<td></td>
</tr>
<tr>
<td>Eutectic liq = cz + cyb</td>
<td>2691</td>
<td>0.85</td>
<td>(21)</td>
<td></td>
</tr>
<tr>
<td>Decomposition of delta yb</td>
<td>1898</td>
<td>0.4</td>
<td>(20)</td>
<td></td>
</tr>
</tbody>
</table>

There is very little thermodynamic data for the Y₂O₃-ZrO₂ system. Belov and Semenov have measured activities in the system at 2773 K (22). For the intermediate compound Zr₃Y₄O₁₂, there are some recent first principles calculations of enthalpies at 0 K (23).

As noted, there is much less data for the Yb₂O₃-ZrO₂ system. The liquidus measurement of Rouanet (21) was used in the assessment. Limited phase boundary measurements in the ZrO₂ rich side by Stubican et al. (20) were used. The single reported invariant point is given in Table II. This system also contains an intermediate compound, Zr₃Yb₄O₁₂, which decomposes at a higher temperature than the analogous Zr₃Y₄O₁₂ compound (20).

The third binary needed is the Y₂O₃-Yb₂O₃ system. The only experimental phase diagram (24) shows a continuous cubic solid solution. Thus this system was simply modeled as an ideal solution.
One study of the Y_2O_3-Yb_2O_3-ZrO_2 system was found (25). The major features of this are continuous tetragonal ZrO_2 with Y_2O_3 and Yb_2O_3; continuous cubic ZrO_2 with Y_2O_3 and Yb_2O_3; continuous cubic (Y,Yb)_2O_3; and a continuous line compound Zr_3(Y,Yb)_4O_12. The presence of the line compound, of course, depends on the disordering temperatures in Table II.

**MODELING**

The first step in a Calphad modeling is to determine the lattice stabilities for all phases. These were taken from the SGTE pure substance database (26), which in turn is from the IVTAN tables (27). These numbers are derived from the phase transition temperature and heats. We have modified the melting point of Y_2O_3 to agree with the melting point measured by Stubican (12). The SGTE database includes a solid/solid and a solid/liquid phase transition for Yb_2O_3, but only the temperatures of transition are given—no heats are reported. Therefore we used the heats for analogous Y_2O_3 transitions.

Two commercial codes were used for these calculations—ThermoCalc* and WinPhad**. Both express the solutions as:

\[ G_m(x,T) = \sum_{i=1}^{2} x_i G_i(T) + RT \sum_{i=1}^{2} x_i \ln x_i + \sum_{j=0}^{\infty} x_i x_2 \sum_{j=0}^{\infty} L(x_i - x_2)^j \]  

The first term in equation (1) is the sum of the Gibbs energies of each component, the second term is the ideal entropy of mixing, and the third term is the excess Gibbs energy. Both codes express the excess Gibbs energy as a Redlich-Kister polynomial (28). In all cases, only the first one or two terms were needed, which corresponds to the regular and sub-regular solution models, respectively.

The solution parameters, were derived from the selected experimental data using the optimizing modules of ThermoCalc and WinPhad. The most prominent features of the phase diagrams were first developed and then the fine features added. The goal is to develop a description of the entire system with a minimum number of parameters. As noted, binaries can be put together to form ternaries. In general a limited number of ternary interaction coefficients are necessary.

**RESULTS**

Figure 1 is the calculated Y_2O_3-ZrO_2 pseudo-binary. In general agreement with experimental data is good. The four invariant points are reproduced as is the disordering temperature for Zr_3Y_4O_12. The (tetragonal + cubic)/cubic phase boundary is x(Y_2O_3) ~0.02 high at the lower temperatures. A better model should help reproduce this feature. Current work focuses on the modeling of this system with the ionic model (29).

* ThermoCalc Software, Stockholm, Sweden
** CompuTherm LLC, Madison, Wisconsin
Figure 2 is the calculated Yb$_2$O$_3$-ZrO$_2$ pseudo-binary. Agreement with the limited experimental data is good. Note the intermediate compound decomposes at a higher temperature than that in the Y$_2$O$_3$-ZrO$_2$ system.

Figure 1. Calculated Y$_2$O$_3$-ZrO$_2$ phase diagram. The data points are the experimental points.

Figure 2. Calculated Yb$_2$O$_3$-ZrO$_2$ pseudo-binary phase diagram. Data points are experimental data.
Figure 3 is the calculated pseudo-ternary section at 1673 K. The major experimental features of this ternary (24) have been reproduced. There are continuous tetragonal ZrO₂, cubic ZrO₂, and cubic (Y₂,Yb)₂O₃ phases. Note also the intermediate line compound, Zr₃(Yb,Y)₄O₁₂, does not extend to the Zr₃Y₃O₁₂ side as this compound has already decomposed at 1673 K (see Figure 1).

SUMMARY AND CONCLUSIONS

The Y₂O₃-ZrO₂, Yb₂O₃-ZrO₂, and Y₂O₃-Yb₂O₃ pseudo binaries have been assessed. Available phase boundary and thermodynamic has been used to derive the solution parameters for a Calphad-type database. The oxide units are used as components. The resultant phase diagrams show good agreement with the experimental data. Then the three binaries have been combined to form a database for the Y₂O₃-Yb₂O₃-ZrO₂ pseudo ternary.

ACKNOWLEDGEMENT

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