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ABSTRACT

A simplified kinetic scheme for Jet-A, and methane fuels with water injection was developed to be used in numerical combustion codes, such as the National Combustor Code (NCC) or even simple FORTRAN codes that are being developed at Glenn. The two time step method is either an initial time averaged value (step one) or an instantaneous value (step two). The switch is based on the water concentration in moles/cc of $1 \times 10^{-20}$. The results presented here results in a correlation that gives the chemical kinetic time as two separate functions. This two step method is used as opposed to a one step time averaged method previously developed to determine the chemical kinetic time with increased accuracy. The first time averaged step is used at the initial times for smaller water concentrations. This gives the average chemical kinetic time as a function of initial overall fuel air ratio, initial water to fuel mass ratio, temperature, and pressure. The second instantaneous step, to be used with higher water concentrations, gives the chemical kinetic time as a function of instantaneous fuel and water mole concentration, pressure and temperature (T4). The simple correlations would then be compared to the turbulent mixing times to determine the limiting properties of the reaction.

The NASA Glenn GLSENS kinetics code calculates the reaction rates and rate constants for each species in a kinetic scheme for finite kinetic rates. These reaction rates were then used to calculate the necessary chemical kinetic times. Chemical kinetic time equations for fuel, carbon monoxide and NOx were obtained for Jet-A fuel and methane with and without water injection to water mass loadings of 2/1 water to fuel.

A similar correlation was also developed using data from NASA’s Chemical Equilibrium Applications (CEA) code to determine the equilibrium concentrations of carbon monoxide and nitrogen oxide as functions of overall equivalence ratio, water to fuel mass ratio, pressure and temperature (T3). The temperature of the gas entering the turbine (T4) was also correlated as a function of the initial combustor temperature (T3), equivalence ratio, water to fuel mass ratio, and pressure.
INTRODUCTION

A simplified kinetic scheme for Jet-A, and methane fuels with water injection was developed to be used in numerical combustion codes, such as the National Combustor Code (NCC) or even simple FORTRAN codes that are being developed at Glenn. The two step kinetic scheme presented here results in a correlation that gives the chemical kinetic time as two separate functions. This two step method is used as opposed to a one step time averaged method (Reference 1) to determine the chemical kinetic time with increased accuracy. The first time averaged step is used at initial with smaller water concentrations of less than $1 \times 10^{-20}$ moles/cc. This gives the average chemical kinetic time as a function of initial overall fuel air ratio, initial water to fuel mass ratio, temperature, and pressure. The second instantaneous step, to be used with higher water concentrations, gives the chemical kinetic time as a function of instantaneous fuel and water mole fractions, pressure and temperature (T4). The simple correlations would then be used with the turbulent mixing times to determine the limiting properties of the reaction.

Water injection into gas turbine engines can be useful in many ways. This includes reduced NOx formation, a lower temperature entering into the turbine (T4) and improving the efficiency and performance of the engine. Water injection has been used in industrial applications, including turbo machinery and diesel engines. Aeronautical applications are still being developed and studied. (Reference 2). The chemical kinetic times for Jet-A fuel and methane with water injection is the focus of the research presented here.

Reaction rates are kinetically limited at low temperatures and mixing limited at very high temperatures. According to the Magnussen model (Reference 3), the fuel oxidation rate will be determined by the maximum of either the chemical kinetic time or the turbulent mixing times of the fuel and air. However, for large numerical solutions it is very tedious to use detailed classical calculations to compare both the kinetic and turbulent mixing times to determine the limits of the reaction. Detailed chemical kinetic schemes are extremely time consuming for two and three dimensional computer calculations for combustors.

Large mechanisms with many intermediate species and very fast radical reactions which cause the equations to be stiff (extremely fast compared to the overall rate, requiring a large number of small time steps), making them very difficult to integrate. Calculations for these extensive mechanisms are repetitive and complex. Using the simplified kinetic scheme developed here to calculate the three chemical kinetic times greatly reduces the amount of time required to compare kinetic reaction times with turbulent mixing times and will reduce the time required to obtain a converged solution. The advantage of extracting the chemical kinetic time for only the species of interest from a detailed computation is that we have only the differential equations of interest to solve, resulting in a much smaller set of equations.

This method is for use in Computational Fluid Dynamics (CFD) calculations where chemical kinetics is important. The current version of NCC requires the user to decide to use either chemical kinetics or the turbulent mixing rates for computing the overall conversion rate. Following detailed conventional methods would not allow for the calculation of both in a reasonable amount of time. The derived method allows for a quick and easy comparison over the complete spectrum of conditions. This scheme is intended for use in numerical combustion codes, but it can also be used as a quick and accurate method to calculate chemical reaction rates.

We have also curve fitted T4 and the equilibrium concentrations of COe, and NOxe using data generated by the NASA Chemical Equilibrium Application code (CEA). Jet-A fuel was represented as $C_{12}H_{23}$, using Krishna Kundu’s twenty three step mechanism (References 4 and

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5). The methane combustion was represented using the GRI-mech version 2.1 mechanism (Reference 10).

GLSENS (Reference 6) was used to integrate the system of equations at constant temperature and pressure, at over 2000 conditions to derive the rate expressions. We have massively correlated the output from GLSENS, the NASA detailed kinetics code, into simple exponential expressions for the chemical kinetic times. It may be reasoned that the presented equations are only as good as the overall mechanism that calculates the data. However, performing the calculations in the conventional manner is also only as good as the mechanism equations and constants that go into them. The first sections of the report, including Mixing and Kinetics, Model Equations, H2O Mass Balance, Determination of the Chemical Kinetic Time, and the kinetic schemes pertain to both Jet-A and methane fuels. The Jet-A equilibrium and chemical kinetic time correlations will then be presented, followed by the methane correlations. Suggestions for using the correlations and a comparison of NOx production data (Reference 2) and the NOx produced by the chemical kinetic time correlations will then be given.

**MIXING AND KINETICS**

The Magnussen model (Reference 3) proposes that the maximum of either the turbulent mixing or the chemical kinetic times will be the limiting factor of a chemical reaction. This model could be explored by numerically calculating both times to compare them. However, using detailed mechanisms, this is a long and tedious process. The calculations would be extremely complicated for the detailed chemical kinetic time. By using the equations presented here to determine the chemical kinetic times and using conventional numerical methods to determine mixing times, the Magnussen model can be applied in a much more convenient way.

\[
\text{Net rate } \frac{\sigma_r}{\sigma_{kinetic}} = \min \left( \frac{A\varepsilon}{k} y_{fuel}, \frac{A\varepsilon y_{oxygen}}{r_f}, \sigma_{kinetic} \right) \quad (1)
\]

Where \( \frac{k}{A\varepsilon} \) equals the turbulent mixing time, \( \tau_m \), with \( k \) being the turbulent kinetic energy, \( \varepsilon \) is the dissipation rate, \( y \) is the mass fraction, and \( r_f \) is the stoichiometric coefficient written on a mass fraction basis. The mixing constant, \( A \), is usually given as 4.0. The factor \( \frac{y_{fuel}}{\sigma_{kinetic}} \) is the chemical kinetic time \( \tau_c \) computed in this report from the correlations presented.

In order to obtain the chemical source term \( \sigma_r \), a comparison is made of the mixing rate, \( \frac{1}{\tau_m} \), and the chemical kinetic rate \( \frac{1}{\tau_c} \), and the lowest rate or the longest time is used in the expression; see Figure 1. This may also be represented by the following relationship:

\[
\tau = \max \left( \tau_m, \tau_c \right) \quad (2)
\]
MODEL EQUATIONS

The following equations can be used to model the chemical system.

\[ C_x H_y + (x/2 + y/4) O_2 \xrightarrow{\tau_{\text{Fuel}}} x \text{ CO} + y/2 \text{ H}_2\text{O} \quad (S1) \]

\[ \text{CO} + \frac{1}{2} \text{O}_2 \xrightarrow{\tau_{\text{CO}}} \text{CO}_2 \quad (S2) \]

\[ \text{N}_2 + \text{O}_2 \xrightarrow{2\tau_{\text{NO}}} 2\text{NO} \quad (S3) \]

The following first order reaction was used to represent the rate of fuel burning. (In this report, \( t \) and \( \tau \) are given in milliseconds, except \( \tau_{\text{NOx}} \) which is in \( \text{gmol} \cdot \text{ccms} \), while concentrations are given in gmoles/cc):

\[ \frac{d\text{Fuel}}{dt} = -\frac{\text{Fuel}}{\tau_{\text{Fuel}}} \quad (3) \]

For a constant \( \tau_{\text{Fuel}} \), the fuel concentration is then represented by a simple exponential decay expression, where \( F_0 \) is the initial fuel concentration.

\[ \text{Fuel} = F_0 e^{-\frac{t}{\tau_{\text{Fuel}}}} \quad (4) \]

The carbon monoxide reaction rate was represented by Equations (5) and (5a). The fuel concentration is multiplied by a factor of 12 because the Jet-A fuel takes the formula C_{12}H_{23}. Equation (6) is the solution to the differential equation showing the CO concentration as a function of initial fuel concentration, CO equilibrium concentration and the chemical kinetic times for fuel and CO.

\[ \frac{d\text{CO}}{dt} = -\frac{\text{(CO} - \text{COeq}) + 12\text{Fuel}}{\tau_{\text{CO}}} + \frac{\text{12Fuel}}{\tau_{\text{Fuel}}} \quad (5) \]

and

\[ \frac{d\text{CO}_2}{dt} = \frac{\text{CO}}{\tau_{\text{CO}}} \quad (5a) \]
\[ CO - CO_{eq} = e^{\frac{-t}{\tau_{CO}}} \left[ CO(t = 0) - CO_{eq} = \frac{12F_0 \tau_{CO}}{\tau_f - \tau_{CO}} \right] + \frac{12F_0 \tau_{CO}}{\tau_f - \tau_{CO}} e^{\frac{-t}{\tau_f}} \]  

(6)

Finally, the nitrogen oxide formation rate, a species important for combustor emissions, was modeled as a simple zero order expression.

\[ \frac{dNOx}{dt} = \frac{1}{\tau_{NOx}} \]  

(7)

or  

\[ NOx = \frac{t}{\tau_{NOx}} \]  

(8)

* \( \tau_{NOx} \) has units of \( \frac{ms \cdot cc}{gmol} \)

**H2O MASS BALANCE**

The inlet mixture contains only fuel, H2O, and air, so the initial mole fractions can be easily calculated using the method of LSENS (Reference 7). This procedure is described in detail below. The mixture is completely specified by fixing the equivalence ratio, Eratio, and the water to fuel ratio, H2OF (weight H2O/weight fuel). We have chosen to keep the weight of the H2O separate from the weight of the fuel so that the stoichiometric fuel/air ratio is always 0.068 for all Jet-A water to fuel ratios and 0.059 for all methane water to fuel ratios. This H2O mass balance was used for both Jet-A and methane fuels. The term MWF can be used to represent the molecular weights of either fuel.

Let \( y_i = \frac{\text{moles } i}{\text{mole mixture}} \) \hspace{1cm} (9); \hspace{1cm} let \( x_i = \frac{\text{moles } i}{\text{moles } O_2 \text{ in air}} \) \hspace{1cm} (10)

For the general chemical equation:

\[ C_{n_c} H_{n_h} O_{n_o} + \frac{4n_c + n_h - 2n_o}{4} O_2 = n_c CO_2 + \frac{n_h}{2} H_2 O \]  

(11)

let \( \phi = \frac{f}{O_2} \) \hspace{1cm} (12) ; \hspace{1cm} let \( y_f = \frac{4\phi}{4n_c + n_h - 2n_o} \) \hspace{1cm} (13)

The above equations are the same as with water injection and \( n_o \) is equal to zero for Jet-A and methane. The sum of the mole fractions of all species in the system is equal to one.
\[ y_f + y_{H_2O} + y_{O_2} + y_{N_2} + y_{Ar} + y_{CO} = 1.0 \]  

(14)

let \( H2OF = \frac{lbs \, H_2O}{lb \, fuel} \); \( H2OM = \frac{mole \, H_2O}{mole \, fuel} = \frac{H2OF \, * \, MWF}{18} \)  

(15)

where MWF for Jet-A is 167 and MWF for methane is 16.

\[ y_{H_2O} = y_f \left( \frac{H2OF \, * \, MWF}{18} \right) \]  

(16)

\[ y_f = \frac{4\phi}{4n_c + n_h + 2n_o} \, y_{O_2} \]  

(17)

Or,

\[ y_{O_2} = \frac{4n_c + n_h - 2n_o}{4\phi \left( 1 + \frac{H2OF \, * \, MWF}{18} \right)} + \frac{4n_c + n_h - 2n_o}{(1 + x_{N_2} + x_{Ar} + x_{CO_2})} \]  

(18)

(See the computer code modifications in Appendix C)

**DETERMINATION OF CHEMICAL KINETIC TIME**

With the approach derived here, a simple direct comparison can be made between the mixing and chemical kinetic times and the minimum rate used for the computation as shown in Figure 1. The integration was performed for 2160 cases shown below for Jet-A and methane fuels with water injection.

<table>
<thead>
<tr>
<th>Table Input</th>
<th>Range</th>
</tr>
</thead>
<tbody>
<tr>
<td><strong>Input Parameter</strong></td>
<td><strong>Range</strong></td>
</tr>
<tr>
<td>Pressure</td>
<td>1 to 40 atmospheres (increments of 10 atm)</td>
</tr>
<tr>
<td>Temperature</td>
<td>1000 to 2500K (increments of 500K)</td>
</tr>
<tr>
<td>Lean Equivalence ratios</td>
<td>0.3 to 1.0 (increments of 0.1)</td>
</tr>
<tr>
<td>Rich Equivalence ratios</td>
<td>1.0 to 2.0 (increments of 0.1)</td>
</tr>
<tr>
<td>Water to fuel mass ratio</td>
<td>0.0 to 2.0 (increments of 0.5)</td>
</tr>
<tr>
<td>Step One Time</td>
<td>(1 \times 10^{-6}) to 2 ms</td>
</tr>
<tr>
<td>Step Two Time</td>
<td>0.05 to 6 ms</td>
</tr>
</tbody>
</table>
Calculations were performed isothermally using GLSENS for each condition over a time of 0 to 6 milliseconds. By computing the progress isothermally, the chemical rate constants were fixed and the chemical kinetic time was determined as a unique value of temperature, pressure and instantaneous mole fractions of fuel and water. GLSENS computes the cumulative rate of reaction for each species from all equations in the mechanism, so it is a simple matter to then compute the chemical kinetic time for each species. For the fuel equation (3) the chemical kinetic time is given as

$$\tau_f = -\frac{Fuel}{\left(\frac{dFuel}{dt}\right)}$$

(19)

This simple calculation was done using additional steps in Subroutine Out2 in the GLSENS code (see Appendix D). Values for the chemical kinetic time were calculated for each concentration at each output time and each set of conditions. For time step 1, the trapezoidal rule (using 1/\(\tau\)) was then used to average the chemical kinetic time to calculate the best value for each set of conditions and the final numbers regressed over the complete set of cases to obtain the final correlation.

A correlation could then be developed that determines the chemical kinetic time as a function of the initial overall cell fuel/air ratio, water to fuel mass ratio, pressure and temperature. The data was correlated using the same method as previously mentioned for the equilibrium equations. Two correlations for each step for each of the three species, one for the lean side and one for the rich side, were obtained. This results in a total of 12 Jet-A correlations and 12 Methane correlations. As OH is formed, the reaction rates get faster and the chemical kinetic time gets smaller. Chemical kinetic time was correlated using two different steps to increase the accuracy of the calculation. Step one is an average chemical kinetic time taken over 2.0 milliseconds. Step two is an instantaneous value that depends on the instantaneous amounts of fuel, water and oxygen. Since we were not tracking the radical concentrations, we were using H2O to indicate the state of the radical species. In other words for the reactions

- \(H_2 + OH = H_2O + H\)
- \(CO + OH = CO_2 + H\)
- \(CH_4 + OH = CH_3 + H_2O\)
- \(N + OH = NO + H\) etc.

The radicals could be correlated with H2O because of the first reaction.

The user should want to switch from step one to step two when the molar concentration of water is greater than \(1 \times 10^{-20}\) moles/cc. Step two can not be used with small concentrations of water because if the value of the water concentration was zero, the entire correlation time would go to zero.
JET-A KINETIC SCHEME FOR STEP 1(AVERAGE) AND STEP 2
(INSTANTANEOUS) METHODS

The following is GLSENS input for the 23 step, 16 species mechanism from Krishna Kundu that was used for the Jet-A calculations. The water and fuel were added as a liquid to the equilibrium program, but they were added as gases to the kinetic program to make up the reacting mixture. We did not alter the mechanism equations to compute the chemical kinetic times for water injection. The additional water took part in the reaction set as the rate equations dictated.

Jet-A Mechanism used in GLSENS

\[
\begin{align*}
&\text{H}_2 + \text{OH} = \text{H}_2\text{O} + \text{H} & 1.17\times10^{11} & 1.1 & 3626. \\
&\text{H}_2 + \text{O} = \text{H} + \text{OH} & 2.50\times10^{15} & 0. & 6000. \\
&\text{H} + \text{O}_2 = \text{O} + \text{OH} & 4.00\times10^{14} & 0. & 18000. \\
&\text{N}_2 + \text{O}_2 > 2.0\text{O} + \text{N}_2 & 1.00\times10^{18} & 0. & 122239. \\
&\text{H}_2 + 2.0\text{O} > \text{O}_2 + \text{H}_2 & 5.00\times10^{17} & 0.5 & 0. \\
&\text{H}_2 + 2.0\text{H} = 2.0\text{H}_2 & 4.00\times10^{20} & -1. & 0. \\
&\text{H} + \text{O}_2 = \text{HO}_2 & 1.00\times10^{15} & -1.1 & 0. \\
&\text{O} + \text{HO}_2 = \text{OH} + \text{O}_2 & 1.50\times10^{13} & 0. & 0. \\
&\text{H} + \text{HO}_2 = \text{H}_2 + \text{O}_2 & 1.50\times10^{13} & 0. & 0. \\
&\text{CO} + \text{OH} = \text{CO}_2 + \text{H} & 4.17\times10^{11} & 0.0 & 1000. \\
&\text{CO} + \text{HO}_2 > \text{CO}_2 + \text{OH} & 5.80\times10^{13} & 0. & 22934. \\
&\text{CH} + \text{O} = \text{CO} + \text{H} & 1.00\times10^{10} & 0.5 & 0. \\
&\text{CH} + \text{NO} = \text{CO} + \text{NH} & 1.00\times10^{11} & 0. & 0. \\
&\text{CH} + \text{HO}_2 > \text{CO}_2 + \text{OH} & 3.00\times10^{12} & 0. & 0. \\
&\text{C}_2\text{H}_2 + \text{O} = \text{CO}_2 + \text{H} & 3.00\times10^{12} & 0. & 49000. \\
&\text{N}_2 + 2.0\text{N} = \text{N}_2 + \text{N}_2 & 1.00\times10^{15} & 0. & 0. \\
&\text{N} + \text{O}_2 = \text{NO} + \text{O} & 6.30\times10^{09} & 1. & 6300. \\
&\text{N} + \text{OH} = \text{NO} + \text{H} & 3.00\times10^{13} & 0. & 0. \\
&\text{NH} + \text{O} = \text{NO} + \text{H} & 1.50\times10^{13} & 0. & 0. \\
&\text{NH} + \text{NO} = \text{N}_2 + \text{OH} & 2.00\times10^{15} & -0.8 & 0. \\
&\text{O} + \text{N}_2 + \text{H}_2 > 2.0\text{NO} + \text{H} + \text{O} & 1.50\times10^{07} & 1. & 45900. \\
&2.0\text{NO} + \text{H} > \text{N}_2 + \text{HO}_2 & 2.50\times10^{10} & 0.16 & 8000. \\
&\text{N}_2 + \text{O} > \text{NO} + \text{N} & 4.75\times10^{10} & 0.29 & 75010. \\
&\text{N} + \text{NO} > \text{N}_2 + \text{O} & 3.00\times10^{12} & 0.2 & 0. \\
&\text{H}_2 + \text{N}_2 > 2.0\text{O} + \text{H}_2 & 1.00\times10^{16} & 0. & 78000. \\
&2.0\text{OH} > 2.0\text{O} + \text{H}_2 & 2.50\times10^{09} & 0. & 30000. \\
&\text{N}_2 + \text{C}_2\text{H}_2 > 6.0\text{C}_2\text{H}_2 & 1.95\times10^{15} & 0. & 0. \\
&\text{N}_2 + \text{C}_2\text{H}_2 > 12.0\text{CH} + 6.0\text{C}_2\text{H}_2 & 2.50\times10^{09} & 0. & 30000. \\
\end{align*}
\]

For example the last three body mechanism step the rate is given by

\[2.5 \times 10^{10} T^{-0.6} \exp \left( -30000/RT \right) N_2^{0.8} C_{12}H_{23}^{3.8}\]

in an irreversible step.

Note the fuel is C_{12}H_{23}. The last two steps are irreversible fuel breakup reactions to CH and C_{2}H_{2}. 

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Note, some reactions are bimolecular and some are trimolecular expressions. The code follows the method of LSENS developed by Radhakrishnan (Reference 7).

**PARTIAL METHANE KINETICS SCHEME**

Whereas Jet-A is broken down in an irreversible molecule breaking step which has a given rate, methane, CH$_4$ is broken down by radicals OH, H, O, etc so the fast reaction is delayed until the radical pool builds up. This forces the modeling scheme to use a two step method to predict the chemical time constants. We have just listed the CH$_4$ reactions to illustrate this process. The complete GRI-mech mechanism is listed in Appendix B. The complicated nature of the methane mechanism makes it very difficult to obtain a correlation with so few variables over a wide range of conditions (so the resulting $R^2$ is low).

CH$_4$ partial mechanism showing free radical attack

<table>
<thead>
<tr>
<th>Reaction</th>
<th>Species</th>
<th>Rate Constant</th>
<th>Reaction Order</th>
<th>Activation Energy</th>
</tr>
</thead>
<tbody>
<tr>
<td>OH + CH$_4$ = CH$_3$ H$_2$O</td>
<td>1.000E+08</td>
<td>1.600</td>
<td>3120.00</td>
<td></td>
</tr>
<tr>
<td>OH + CO = H CO$_2$</td>
<td>4.760E+07</td>
<td>1.228</td>
<td>70.00</td>
<td></td>
</tr>
<tr>
<td>N + OH = NO H</td>
<td>7.333E+13</td>
<td>0.000</td>
<td>1120.00</td>
<td></td>
</tr>
<tr>
<td>2.0 O = O$_2$ M</td>
<td>1.200E+17</td>
<td>-1.000</td>
<td>0.00</td>
<td></td>
</tr>
<tr>
<td>H + H = H$_2$ M</td>
<td>1.000E+18</td>
<td>-1.000</td>
<td>0.00</td>
<td></td>
</tr>
<tr>
<td>O + H$_2$ = H OH</td>
<td>5.000E+04</td>
<td>2.670</td>
<td>6290.00</td>
<td></td>
</tr>
<tr>
<td>H + O$_2$ = O OH</td>
<td>8.300E+13</td>
<td>0.000</td>
<td>14413.00</td>
<td></td>
</tr>
<tr>
<td>H + HO$_2$ = O$_2$ H$_2$</td>
<td>2.800E+13</td>
<td>0.000</td>
<td>1068.0</td>
<td></td>
</tr>
<tr>
<td>OH + H$_2$ = H H$_2$O</td>
<td>2.160E+08</td>
<td>1.510</td>
<td>3430.00</td>
<td></td>
</tr>
<tr>
<td>O + CH$_4$ = OH CH$_3$</td>
<td>1.020E+09</td>
<td>1.500</td>
<td>8600.00</td>
<td></td>
</tr>
<tr>
<td>OH + CH$_2$ = CH H$_2$O</td>
<td>1.130E+07</td>
<td>2.000</td>
<td>3000.00</td>
<td></td>
</tr>
<tr>
<td>OH + CH$_3$ = CH$_3$OH M</td>
<td>6.300E+13</td>
<td>0.000</td>
<td>0.00</td>
<td></td>
</tr>
<tr>
<td>CH + CH$_4$ = H C$_2$H$_4$</td>
<td>6.000E+13</td>
<td>0.000</td>
<td>0.00</td>
<td></td>
</tr>
<tr>
<td>CH$_2$ + CH$_4$ = 2.0CH$_3$</td>
<td>2.460E+06</td>
<td>2.000</td>
<td>8270.00</td>
<td></td>
</tr>
<tr>
<td>CH$_3$ + H$_2$O = HCO CH$_4$</td>
<td>3.320E+03</td>
<td>2.810</td>
<td>5860.00</td>
<td></td>
</tr>
<tr>
<td>CH$_3$ + CH$_3$OH = CH$_2$OH CH$_4$</td>
<td>3.000E+07</td>
<td>1.500</td>
<td>9940.00</td>
<td></td>
</tr>
<tr>
<td>CH$_3$ + CH$_3$OH = CH$_3$O CH$_4$</td>
<td>1.000E+07</td>
<td>1.500</td>
<td>9940.00</td>
<td></td>
</tr>
<tr>
<td>CH$_3$ + C$_2$H$_4$ = C$_2$H$_3$ CH$_4$</td>
<td>2.270E+05</td>
<td>2.000</td>
<td>9200.00</td>
<td></td>
</tr>
<tr>
<td>CH$_3$ + C$_2$H$_6$ = C$_2$H$_5$ CH$_4$</td>
<td>6.140E+06</td>
<td>1.740</td>
<td>10450.00</td>
<td></td>
</tr>
<tr>
<td>N + O$_2$ = NO O</td>
<td>2.650E+12</td>
<td>0.000</td>
<td>6400.00</td>
<td></td>
</tr>
<tr>
<td>NO + O = NO$_2$ M</td>
<td>1.060E+20</td>
<td>-1.410</td>
<td>0.00</td>
<td></td>
</tr>
</tbody>
</table>

**JET-A EQUILIBRIUM CORRELATIONS WITH AND WITHOUT WATER INJECTION**

Equilibrium correlations were generated by using Microsoft Excel to perform a multivariate linear regression on the large data set generated by the CEA program of Reference 8. (A detailed procedure describing the regression used for both equilibrium and finite rate chemical times can be found in Appendix A). CEA has a plot f option for direct tabulation of the output data, for ‘f/a, P, T H$_2$O, CO, NO’. Although the equilibrium correlations were not usually used in the calculation of the chemical kinetic times, we feel that these equations could still be very useful for other calculations. Table 1 shows the Jet-A equilibrium correlations for
T4, CO and NOx for both the lean and rich cases. Figures 2 is a parity plot showing the strength of the lean Jet-A CO equilibrium correlation. This plot shows a minimal amount of scatter, mostly at an equivalence ratio of 1.0, indicating a strong correlation ($R^2$ values greater than 0.9). Note that the units of $CO_{eq}$ and $NOx_{eq}$ are moles/cc. This parity plot is typical of all the variables because of the high $R^2$ values obtained. This was a correlation for all of the range of independent variables in the Table Input.

**Table 1. EQUILIBRIUM Lean Jet-A with Water Injection Correlations**

<table>
<thead>
<tr>
<th>Species</th>
<th>Lean ($f/a \leq 0.068$)</th>
<th>R-squared</th>
</tr>
</thead>
<tbody>
<tr>
<td>T4</td>
<td>$T_4 = 1725 , (T^3)^{0.241} \left( \frac{f}{a} \right)^{0.442} (1 + h2a/f)^{(-0.151)} P^{0.00301}$</td>
<td>0.949</td>
</tr>
<tr>
<td>CO</td>
<td>$CO_{eq} = 226(\frac{f}{a})^{0.222} (1 + h2a/f)^{0.0658} P^{0.524} \exp\left[\frac{-31647}{T}\right]$</td>
<td>0.995</td>
</tr>
<tr>
<td>NOx</td>
<td>$NOx_{eq} = 2.65e^{-8} (\frac{f}{a})^{(-1.52)} (1 + h2a/f)^{(-0.133)} P^{0.980} \exp\left[\frac{-9953}{T}\right]$</td>
<td>0.958</td>
</tr>
</tbody>
</table>

**Table 2. EQUILIBRIUM Rich Jet-A with Water Injection Correlations**

<table>
<thead>
<tr>
<th>Species</th>
<th>Rich ($f/a &gt; 0.068$)</th>
<th>R-squared</th>
</tr>
</thead>
<tbody>
<tr>
<td>T4</td>
<td>$T_4 = 163(T^3)^{0.186} \left( \frac{f}{a} \right)^{(-0.613)} (1 + h2a/f)^{(-0.296)} P^{0.00231}$</td>
<td>0.959</td>
</tr>
<tr>
<td>CO</td>
<td>$CO_{eq} = 3.85e^{-4} (\frac{f}{a})^{2.86} (1 + h2a/f)^{(-0.369)} P^{0.995} \exp\left[\frac{185}{T}\right]$</td>
<td>0.990</td>
</tr>
<tr>
<td>NOx</td>
<td>$NOx_{eq} = 1.80e^{-8}(\frac{f}{a})^{(-0.608)} (1 + h2a/f)^{0.0418} P^{0.530} \exp\left[\frac{-38952}{T}\right]$</td>
<td>0.993</td>
</tr>
</tbody>
</table>

**CHEMICAL KINETIC TIMES FOR JET-A WITH AND WITHOUT WATER INJECTION**

**Step One Equations for Jet-A**

The following form of equation was used for the Jet-A step one correlations:

$$\tau = A(P)^a \left( \frac{f}{a} \right)^b (1 + \frac{h2a}{fuel})^c \exp \frac{D}{T}$$ (26)
where \( \tau \) is the chemical kinetic time in milliseconds, \( P \) is pressure in atm, \( f/a \) is the initial or overall mass fuel air ratio as in reference 1, \( \frac{h_{2o}}{fuel} \) is the initial mass water to fuel ratio and \( T \) is the temperature in Kelvin. The correlation is switched to Step Two when the water molar concentration is greater than \( 1 \times 10^{-20} \).

The initial or overall fuel air ratio can be determined as follows:

\[
f / a = \left( fuel + \frac{(CO + CO_2)}{12} \right) * 167 / \frac{0.79}{28N_2}
\]

(27)

The coefficients for each of the parameters in the correlations may be found in Table 3. R-square values have been included to demonstrate the strength of the correlation. R-squared is a measure of the error that the model accounts for; an R-squared value of one is ideal.

<table>
<thead>
<tr>
<th>Component</th>
<th>Rich or lean</th>
<th>A</th>
<th>a</th>
<th>D</th>
<th>b</th>
<th>c</th>
<th>R-squared</th>
</tr>
</thead>
<tbody>
<tr>
<td>Fuel</td>
<td>Lean</td>
<td>7.47x10^-5</td>
<td>-0.60</td>
<td>14202</td>
<td>0.238</td>
<td>0.0712</td>
<td>1.00</td>
</tr>
<tr>
<td>CO</td>
<td>Lean</td>
<td>7.13x10^-2</td>
<td>-0.758</td>
<td>9295</td>
<td>-0.314</td>
<td>0.159</td>
<td>0.933</td>
</tr>
<tr>
<td>NOx</td>
<td>Lean</td>
<td>1.00x10^-8</td>
<td>-1.30</td>
<td>26139</td>
<td>0.110</td>
<td>1.30</td>
<td>0.994</td>
</tr>
<tr>
<td>Fuel</td>
<td>Rich</td>
<td>8.19x10^-5</td>
<td>-0.60</td>
<td>14206</td>
<td>0.296</td>
<td>0.153</td>
<td>1.00</td>
</tr>
<tr>
<td>CO</td>
<td>Rich</td>
<td>1.39x10^-3</td>
<td>-0.882</td>
<td>6803</td>
<td>-0.222</td>
<td>-0.328</td>
<td>0.988</td>
</tr>
<tr>
<td>NOx</td>
<td>Rich</td>
<td>93.2</td>
<td>-1.67</td>
<td>27755</td>
<td>-0.0582</td>
<td>0.529</td>
<td>0.983</td>
</tr>
</tbody>
</table>

(28)  (29)  (30)  (31)  (32)  (33)

All of the results presented here are correlated over the complete conditions listed in Table Input. Parity plots for the lean step one correlations have been created and may be found in Figures 3-5. The x-axis contains values for the chemical kinetic time generated by the full mechanism GLSENS at each condition. The y-axis contains values calculated using the chemical kinetic time correlations above at the same set of conditions. This demonstrates how close the calculated value is to the expected value and is a good measurement of the strength of the correlations. Figure 6 shows an increase in the step one chemical kinetic time with a higher water to fuel ratio. Data from the previous Jet-A chemical kinetic time correlation without water injection has also been included on this plot for comparison. In order to compare the correlations, we took various values of \( f/a, p, \) and \( T \) and obtained predictions for the previous correlation (Ref. 1 as “No Water Injection”). Then these prediction are plotted using the overall factor found here (x axis). Note that the previous work gave a much faster time (averaged over a total time as high as 10 milliseconds and accounted for the lower \( \tau \)). This difference in results caused us to switch to an instantaneous correlation (Step Two) for longer times.

**Step Two Equations for Jet-A**

The following form of equation was used for the Jet-A step two fuel and NOx correlations. This form of the equation produced the best fit:
\[ \tau = A(P)^a (c_{fuel})^b (c_{o2})^c (c_{H2O})^d \exp \frac{E}{T} \]  

(34)

and for the Jet-A step two CO correlation:

\[ \tau = A(P)^a (c_{fuel})^b (c_{CO})^c (c_{o2})^d (c_{H2O})^e \exp \frac{F}{T} \]  

(35)

where \( P \) is pressure in atm, \( c_{fuel} \) is the instantaneous molar concentration of fuel, \( c_{CO} \) is the instantaneous molar concentration of CO. We added CO as a parameter in order to attempt to raise the \( R^2 \) value for the CO correlation. We have preceded the symbol with a c to indicate molar concentration was used and a y for mole fraction. The CO2 is the instantaneous molar concentration of O2, \( c_{H2O} \) is the instantaneous molar concentration of water, and \( T \) is the temperature in Kelvin. We have correlated to only the major species hoping that H2O will track the minor species (OH, H, O, etc.) to allow good overall correlation and easy use of the equations.

The coefficients for each parameter are given in Table 4. Parity plots for the step two lean Jet-A correlations can be found in Figures 7-9. These figures show a minimal amount of scattering for the fuel and NOx, which is consistent with the high R-squared values as seen in Table 4. However, the CO plot shows considerably more scattering with an R-squared value of 0.578 for the lean case and 0.389 for the rich case.

### Table 4. Step Two Jet-A With Water Injection Chemical Kinetic Time Correlations

<table>
<thead>
<tr>
<th>Component</th>
<th>Rich or Lean</th>
<th>A</th>
<th>Pressure</th>
<th>1/T</th>
<th>cfuel</th>
<th>cCO</th>
<th>cO2</th>
<th>cH2O</th>
<th>R-squared</th>
</tr>
</thead>
<tbody>
<tr>
<td>Fuel</td>
<td>Lean</td>
<td>7.31E-06</td>
<td>-0.173</td>
<td>12412</td>
<td>0.0792</td>
<td>-</td>
<td>-0.164</td>
<td>-0.115</td>
<td>0.927</td>
</tr>
<tr>
<td>CO</td>
<td>Lean</td>
<td>9.28E-07</td>
<td>-0.164</td>
<td>8893</td>
<td>-0.15</td>
<td>0.268</td>
<td>-0.549</td>
<td>-0.0588</td>
<td>0.578</td>
</tr>
<tr>
<td>NOx</td>
<td>Lean</td>
<td>2.67E-03</td>
<td>-0.628</td>
<td>28071</td>
<td>-0.186</td>
<td>-</td>
<td>-0.558</td>
<td>0.0458</td>
<td>0.93</td>
</tr>
<tr>
<td>Fuel</td>
<td>Rich</td>
<td>1.35E-04</td>
<td>-0.352</td>
<td>12962</td>
<td>0.0147</td>
<td>-</td>
<td>-0.0743</td>
<td>-0.0373</td>
<td>0.953</td>
</tr>
<tr>
<td>CO</td>
<td>Rich</td>
<td>0.373</td>
<td>-0.422</td>
<td>4387</td>
<td>-0.287</td>
<td>0.206</td>
<td>0.227</td>
<td>0.115</td>
<td>0.389</td>
</tr>
<tr>
<td>NOx</td>
<td>Rich</td>
<td>28.9</td>
<td>-0.00805</td>
<td>19595</td>
<td>-0.117</td>
<td>-</td>
<td>-0.158</td>
<td>-0.16</td>
<td>0.438</td>
</tr>
</tbody>
</table>

### METHANE EQUILIBRIUM CORRELATIONS WITH AND WITHOUT WATER INJECTION

The rich and lean Methane equilibrium correlations can be found in Tables 5 and 6, respectively. A parity plot for the rich Methane CO equilibrium correlation can be found in Figure 10.
Table 5. EQUILIBRIUM Lean Methane With Water Injection Correlations

<table>
<thead>
<tr>
<th>Species</th>
<th>Lean ((f/a \leq 0.058))</th>
<th>R-Squared</th>
</tr>
</thead>
<tbody>
<tr>
<td>T4</td>
<td>(T4 = 1565(f/a)^{0.435}(1 + \frac{h2o}{f})^{-0.069}(T3)^{0.258}(P)^{0.00284})</td>
<td>0.944 (42)</td>
</tr>
<tr>
<td>CO</td>
<td>(COeq = 3.37 \times 10^6 (f/a)^{1.90}(1 + \frac{h2o}{f})^{0.00705}(P)^{-0.477} \exp\left(\frac{33388}{T}\right))</td>
<td>0.998 (43)</td>
</tr>
<tr>
<td>NO(_x)</td>
<td>(NOxeq = 2.44 \times 10^{-2}(f/a)^{-1.13}(1 + \frac{h2o}{f})^{-1.07}(P)^{-0.017} \exp\left(\frac{-11415}{T}\right))</td>
<td>0.953 (44)</td>
</tr>
</tbody>
</table>

Table 6. EQUILIBRIUM Rich Methane With Water Injection Correlations

<table>
<thead>
<tr>
<th>Species</th>
<th>Rich ((f/a &gt; 0.058))</th>
<th>R-Squared</th>
</tr>
</thead>
<tbody>
<tr>
<td>T4</td>
<td>(T4 = 177(f/a)^{-0.559}(1 + \frac{h2o}{f})^{-0.146}(T3)^{0.176}(P)^{0.00167})</td>
<td>0.976 (45)</td>
</tr>
<tr>
<td>CO</td>
<td>(COeq = 72.2(f/a)^{2.55}(1 + \frac{h2o}{f})^{-0.358}(P)^{-0.00393} \exp\left(\frac{-31365}{T}\right))</td>
<td>0.911 (46)</td>
</tr>
<tr>
<td>NO(_x)</td>
<td>(NOxeq = 0.218(f/a)^{-4.43}(1 + \frac{h2o}{f})^{0.427}(P)^{-0.493} \exp\left(\frac{-41408}{T}\right))</td>
<td>0.998 (47)</td>
</tr>
</tbody>
</table>

CHEMICAL KINETIC TIMES FOR METHANE WITH AND WITHOUT WATER INJECTION

Step One Equations for Methane

The step one Methane chemical kinetic time correlations are of the following form:

\[
\tau = A(P)^a (f/a)^b (1 + \frac{h2o}{fuel})^c \exp\left[\frac{D}{T}\right]
\]

(48)

where \(P\) is pressure in atm, \(f/a\) is the initial or overall fuel air ratio, \(\frac{h2o}{fuel}\) is the initial water to fuel ratio and \(T\) is the temperature in Kelvin.

The initial or overall fuel air ratio can be determined as follows:

\[
f/a = (fuel + (CO + CO_2)) \times 16/\frac{0.79}{28N_2}
\]

(49)
A summary of these correlations can be found in Table 7. Parity plots for the lean step one methane with water injection correlations can be found in Figures 11-13. These parity plots show minimal scatter, which is consistent with the high R-squared values of the lean correlations. Figure 14 shows an increase in step one chemical kinetic time with a higher water to fuel ratio.

<table>
<thead>
<tr>
<th>Component</th>
<th>Rich or Lean</th>
<th>A</th>
<th>a</th>
<th>D</th>
<th>b</th>
<th>c</th>
<th>R-squared</th>
</tr>
</thead>
<tbody>
<tr>
<td>Fuel</td>
<td>Lean</td>
<td>2.09x10^{-4}</td>
<td>-1.07</td>
<td>22625</td>
<td>0.222</td>
<td>0.0675</td>
<td>0.996</td>
</tr>
<tr>
<td>Co</td>
<td>Lean</td>
<td>9.99x10^{-4}</td>
<td>-1.00</td>
<td>2434</td>
<td>0.101</td>
<td>0.0959</td>
<td>0.999</td>
</tr>
<tr>
<td>NO_x</td>
<td>Lean</td>
<td>29395</td>
<td>-2.11</td>
<td>34859</td>
<td>0.0315</td>
<td>0.329</td>
<td>0.996</td>
</tr>
<tr>
<td>Fuel</td>
<td>Rich</td>
<td>0.0274</td>
<td>-0.328</td>
<td>14216</td>
<td>0.111</td>
<td>-0.829</td>
<td>0.75</td>
</tr>
<tr>
<td>Co</td>
<td>Rich</td>
<td>1.30x10^{-3}</td>
<td>-1.00</td>
<td>2433</td>
<td>0.215</td>
<td>0.195</td>
<td>0.999</td>
</tr>
<tr>
<td>NO_x</td>
<td>Rich</td>
<td>43928</td>
<td>-2.00</td>
<td>32649</td>
<td>0.284</td>
<td>0.641</td>
<td>0.999</td>
</tr>
</tbody>
</table>

### Step Two Equations for Methane

The following form of equation was used for all lean Methane step two correlations:

\[
\tau = A(P)^a (cfuel)^b (ch2o)^c (1 + \frac{ch2o}{cfuel})^d \exp\left[\frac{E}{T}\right]
\]  \hspace{1cm} (56)

The following form of equation was used for the rich Methane step two correlations for fuel and NO_x:

\[
\tau = A(P)^a (cfuel)^b (co2)^c (ch2o)^d \exp\frac{E}{T}
\]  \hspace{1cm} (57)

and for the rich Methane step two CO correlation:

\[
\tau = A(P)^a (cfuel)^b (cco)^c (co2)^d (ch2o)^e \exp\frac{F}{T}
\]  \hspace{1cm} (58)

Table 8 provides a summary of these correlations, and lean parity plots can be found in Figures 15-17.
Table 8. Step Two Methane With Water Injection Chemical Kinetic Time Correlations

<table>
<thead>
<tr>
<th>Component</th>
<th>A</th>
<th>Pressure</th>
<th>1/T</th>
<th>cfuel</th>
<th>cCO</th>
<th>cO2</th>
<th>cH2O</th>
<th>(1 + \frac{ch_{2}O}{cfuel})</th>
<th>R-squared</th>
</tr>
</thead>
<tbody>
<tr>
<td>Fuel (lean)</td>
<td>3.50x10^{-9}</td>
<td>-0.0713</td>
<td>14149</td>
<td>-0.307</td>
<td>-</td>
<td>-</td>
<td>-0.465</td>
<td>-0.221</td>
<td>0.416</td>
</tr>
<tr>
<td>CO (lean)</td>
<td>3.07x10^{-3}</td>
<td>-1.09</td>
<td>2291</td>
<td>0.0957</td>
<td>-</td>
<td>-</td>
<td>-9.82x10^{-4}</td>
<td>0.131</td>
<td>0.999</td>
</tr>
<tr>
<td>NOx (lean)</td>
<td>1.48x10^{0}</td>
<td>-2.28</td>
<td>37437</td>
<td>0.239</td>
<td>-</td>
<td>-</td>
<td>-0.0206</td>
<td>0.778</td>
<td>0.986</td>
</tr>
<tr>
<td>Fuel (rich)</td>
<td>9.99x10^{-11}</td>
<td>-0.761</td>
<td>19950</td>
<td>0.050</td>
<td>-</td>
<td>-0.549</td>
<td>-0.350</td>
<td>-</td>
<td>0.483</td>
</tr>
<tr>
<td>CO (rich)</td>
<td>1.64E-10</td>
<td>0.0722</td>
<td>7880</td>
<td>-0.261</td>
<td>0.0568</td>
<td>-0.636</td>
<td>-0.00341</td>
<td>-</td>
<td>0.816</td>
</tr>
<tr>
<td>NOx (rich)</td>
<td>1.67E+06</td>
<td>-2.45</td>
<td>24117</td>
<td>0.05</td>
<td>-</td>
<td>0.119</td>
<td>-0.281</td>
<td>-</td>
<td>0.792</td>
</tr>
</tbody>
</table>

Two Step Methane Chemical Kinetic Times Without Water Injection

The step two methane chemical kinetic times were also correlated without water injection. The previous correlations are a massive regression over all water injection values. If one were not using water injection, the relations would be useful, but the following relations resulted in a higher R^2 factor. The equations are of the same form as the methane step two correlations with water injection. The results can be found in Table 9:

Table 9. Step Two Methane Without Water Injection Chemical Kinetic Time Correlations

<table>
<thead>
<tr>
<th>Component</th>
<th>A</th>
<th>Pressure</th>
<th>1/T</th>
<th>cfuel</th>
<th>cCO</th>
<th>cO2</th>
<th>cH2O</th>
<th>(1 + \frac{ch_{2}O}{cfuel})</th>
<th>R-squared</th>
</tr>
</thead>
<tbody>
<tr>
<td>Fuel (lean)</td>
<td>1.86x10^{9}</td>
<td>0.479</td>
<td>13446</td>
<td>0.182</td>
<td>-</td>
<td>0.0528</td>
<td>-0.704</td>
<td>0.349</td>
<td>0.767</td>
</tr>
<tr>
<td>CO (lean)</td>
<td>1.43x10^{-3}</td>
<td>-1.04</td>
<td>2371</td>
<td>0.0461</td>
<td>-</td>
<td>-</td>
<td>-5.7x10^{-4}</td>
<td>-</td>
<td>0.999</td>
</tr>
<tr>
<td>NOx (lean)</td>
<td>10.0</td>
<td>-1.92</td>
<td>38229</td>
<td>0.0389</td>
<td>-</td>
<td>-</td>
<td>-0.200</td>
<td>-</td>
<td>0.992</td>
</tr>
<tr>
<td>Fuel (rich)</td>
<td>2.02x10^{-14}</td>
<td>0.498</td>
<td>14860</td>
<td>0.252</td>
<td>-</td>
<td>-0.776</td>
<td>-0.756</td>
<td>0.254</td>
<td>0.995</td>
</tr>
<tr>
<td>CO (rich)</td>
<td>6.73x10^{-13}</td>
<td>0.244</td>
<td>8229</td>
<td>-0.274</td>
<td>-0.237</td>
<td>-0.802</td>
<td>0.217</td>
<td>-</td>
<td>0.806</td>
</tr>
<tr>
<td>NOx (rich)</td>
<td>0.0574</td>
<td>-1.22</td>
<td>26010</td>
<td>-0.103</td>
<td>-</td>
<td>-0.0583</td>
<td>-0.617</td>
<td>-</td>
<td>0.959</td>
</tr>
</tbody>
</table>

Combined Lean and Rich Step Two Methane Times

In this case the program would not have to choose between the rich and lean fuel/air zones, but could use the correlation directly. The lean and rich step two methane chemical kinetic times were also combined into one large data set and correlated. The results of this correlation can be found in Table 10. The fuel correlation is in the following form:
\[ \tau = A(P)^a (c_{\text{fuel}})^b (c_{\text{O}_2})^c (c_{\text{H}_2\text{O}})^d \exp \frac{E}{T} \]  

(71)

while the CO and NO\textsubscript{x} correlations are modeled by the following form:

\[ \tau = A(P)^e (c_{\text{fuel}})^f (c_{\text{H}_2\text{O}})^g (1 + \frac{c_{\text{H}_2\text{O}}}{c_{\text{fuel}}}) \exp \frac{E}{T} \]  

(72)

<table>
<thead>
<tr>
<th>Component</th>
<th>A</th>
<th>Pressure</th>
<th>1/T</th>
<th>c\textsubscript{fuel}</th>
<th>c\textsubscript{O}_2</th>
<th>c\textsubscript{H}_2\text{O}</th>
<th>1+ \frac{c_{\text{H}<em>2\text{O}}}{c</em>{\text{fuel}}}</th>
<th>R-squared</th>
</tr>
</thead>
<tbody>
<tr>
<td>Fuel</td>
<td>6.20\times10^{-9}</td>
<td>-0.290</td>
<td>14259</td>
<td>-0.0532</td>
<td>-0.332</td>
<td>-0.429</td>
<td>-</td>
<td>0.433</td>
</tr>
<tr>
<td>CO</td>
<td>1.18\times10^{26}</td>
<td>-5.37</td>
<td>-1877</td>
<td>4.68</td>
<td>-0.106</td>
<td>-0.106</td>
<td>4.90</td>
<td>0.627</td>
</tr>
<tr>
<td>NO\textsubscript{x}</td>
<td>2.42\times10^{-14}</td>
<td>0.585</td>
<td>37951</td>
<td>-2.92</td>
<td>-0.0539</td>
<td>-2.98</td>
<td>-2.98</td>
<td>0.879</td>
</tr>
</tbody>
</table>

COMPARISON TO NO\textsubscript{x} DATA WITH WATER INJECTION

An important result of water injection into Jet-A and methane fuels is the reduction of NO\textsubscript{x} formation. The NO\textsubscript{x} concentrations calculated using the chemical kinetic time correlations were compared to water injection data presented in reference 2. This reduction is with constant T4. So the equations from Table 1 Lean T4 is solved for the increased f/a with an increase in h\textsubscript{2}O / f , then the NO\textsubscript{x} value Table 3 Step One Lean is solved for the new NO\textsubscript{x} value. Figures 18 and 19 show the ratio of NO\textsubscript{x} with water injection to NO\textsubscript{x} without water injection versus weight fraction ratio of water to air for Jet-A and methane respectively from the correlations, at a fuel air ratio of 0.05 as shown in reference 2. The Jet-A predicted data behaves very similarly to the given experimental data while there is more of a difference in the methane prediction. Although the temperature is constant and the kinetic mechanism remains the same there is a change in kinetics rate due to water dilution. The residence time of the combustor is decreased because of the increased throughput with water addition (t corrected results) and the molecular weight of the mixture is decreased with water addition (T constant).

TANKS IN SERIES MODEL FOR PREDICTING EMISSIONS

A tanks-in-series Fortran program was developed to simulate CO and NO\textsubscript{x} production, so that it may be compared to the water injection data in reference 9. This model accounts for the increase in fuel/air ratio in the initial fuel injection and mixing process. Figure 20-23 show the results of this comparison.
**SUGGESTIONS FOR USING THE REDUCED EQUATIONS**

In this section we show how our correlations might be used in a typical kinetics program to calculate the amount of reaction.

The fuel equations that we use are:

\[
\frac{d\text{Fuel}}{dt} = -\frac{\text{Fuel}}{\tau_{\text{Fuel}}} \tag{76}
\]
where for step one:

\[
\tau_{\text{Fuel(lean)}} = 7.47 \times 10^{-5} \left(\frac{f}{a}\right)_o^{0.238} (P)^{-0.60} \left(1 + \frac{h2o}{fuel}\right)^{0.0712} \exp \frac{14202}{T}
\]  

(77)

Equation (76) is the “net” rate, so one does not have to compute the reverse rate. Then

\[
\frac{dT}{dt} = -\frac{F \exp \frac{-14202}{T^4}}{7.47 \times 10^{-5} \left(\frac{f}{a}\right)_o^{0.238} (P)^{-0.60} \left(1 + \frac{h2o}{fuel}\right)^{0.0712}} \Rightarrow -\frac{F}{C} e^{-\frac{E}{RT}}
\]  

(78)

The denominator C is a constant since we take the initial conditions as being held constant throughout the reaction and one can treat the T4 constant as an activation energy (14202×1.987) = 28219.

One can relate \((f/a)_o\) to \(F\) of a mixture by:

\[
F_o = C_{12} H_{23} + \frac{1}{12} (CO + CO_2) \frac{\text{moles}}{\text{cc}}
\]  

(79)

\[
\rho = \frac{P}{82.056 T} \frac{\text{moles}}{\text{cc mix}}
\]  

(80)

\[
x_{F_o} = \frac{F_o}{\rho} = \frac{f gm/167}{fgm/167 + A_{gm}/29 + H2O/18F} = \frac{(f/a)_o}{F} + \frac{5.76 + 9.28 H2O/F}{(f/a)_o}
\]  

(81)

Note, \(gm\) is the grams of component. If some fuel had reacted one can use Equation (81) to find \(f/a_o\);

\[
(f/a)_o = \frac{5.76 F/\rho}{1 - F/\rho (1 + 9.28 H2O/F)}
\]  

(82)

Given Equation (78) above

\[
\frac{dT}{dt} \Rightarrow -\frac{F}{C} e^{-\frac{E}{RT}} \text{ which is directly usable by an Arrhenius kinetics code}
\]  

(83)

Then the mass balance equation is
To avoid fractional coefficients, one could use any number of methods, such as:

\[ C_{12}H_{23} + (6 + \frac{23}{4})O_2 = 12CO + \frac{23}{2}H2O \] (84)

For CO

\[ \frac{dCO}{dt} = -\frac{CO}{\tau_{CO}} \] (86)

And for NO

\[ \frac{dNO}{dt} = \frac{1}{\tau_{NO}} = \frac{1}{C'} e^{-\frac{26139}{T}} \] (87)

\[ N_2 + O_2 = 2NO \] (88)

So

\[ \frac{dN_2}{dt} = \frac{1}{2} \frac{dNO}{dt} = \frac{1}{2C'} e^{-\frac{26139}{T}} \] (89)

This completes the equations that one would need to simulate the chemical kinetics. In our model we have only three rates and seven species, therefore the step size can be large and the rates are no longer coupled. This should be very fast.

**SIMULATION OF GLSENS WITH THE TWO STEP MODEL WITHOUT WATER INJECTION**

In this example we compare directly the new two time step method with the complete kinetic mechanism for an arbitrary constant temperature of 2000K and at a constant pressure of two atmospheres. The initial fuel/air ratio was taken as 0.0526 with zero water injection. We used a simple Fortran code with a Newton integration method given in Appendix G. The results are shown in Figures 24 and 25. The results from methane was one reason for creating the two step method with step two trying to follow the instantaneous reaction times though correlation with H2O as the tracking species. Step 2 is very important for prediction of combustion stability and blowout for methane.

Figure 24 shows the rapid decreases in the chemical reaction time. This model is not as fast as the full kinetics version so the concentration decrease (Figure 25) lags the full kinetics version as well. The kinetics calculations in comparison the mixing times from (1.e-5 to 1.e-3 seconds) maybe much more relevant and accurate.
CONCLUSIONS

Much work still needs to be done to explore possible benefits and detriments of water injection in combustion. A simplified kinetic scheme for Jet-A and methane fuels with water injection resulted in a two time step correlation that calculates chemical kinetic times for fuel, CO, and NOx. These chemical kinetic time equations can then be used in a numerical combustor code to compare the chemical kinetic time with the turbulent mixing time. Strong step one Jet-A correlations were developed ($R^2 > 0.9$). The Jet-A step two correlations for CO and NOx are slightly weaker, but still thought to be effective. The Methane step one correlations were all very strong, while the rich step two correlations had considerably smaller R-squared values. However, because we are trying to correlate so many values over a wide range of conditions, we will accept a small amount of scatter. These twenty four equations are believed to be extremely useful in the comparison of kinetic reaction and turbulent mixing times and in the computation of kinetic rate results.
APPENDIX A

MULTIPLE LINEAR REGRESSION WITH EXCEL

Performing Multiple Linear Regression on a Logarithmic Equation

This regression technique may be used to develop a correlation between a dependent variable and one or more independent variables. First the equation to be used must be linearized. An example of an exponential equation used here is shown below.

\[ A = BC^c D^d \exp \left( \frac{e}{T} \right) \]  \hspace{1cm} \text{(A1)}

\[ \ln(A) = \ln(B) + c \ln(C) + d \ln(D) + \frac{e}{T} \]  \hspace{1cm} \text{(A2)}

Columns of data containing the independent variables (natural log of C, natural log of D, 1/T,) and the independent variable (natural log of A) were contained in an Excel spreadsheet. (It is easiest to have the independent variables adjacent to each other, followed by the dependent variable.)

The multiple variable regression analysis is located in the Data Analysis Toolpak. The Data Analysis Toolpak must be added into the spreadsheet if it is not already running in Excel. In order to add it, select the ‘Add ins’ button from the Tools menu. Click on the Analysis Toolpak option and click OK to accept this choice. Then choose ‘Data Analysis’ from the Tools menu and double click on ‘regression’. Click on the ‘Input Y Range’ box and highlight the column that contains the logarithm of the dependent variable and press return. Click on the ‘Input X Range’ box and highlight the columns containing all of the independent variables. (In this case \(\ln(C)\), \(\ln(D)\) and \(1/T\)). Press OK to begin the regression. The regression data will be contained in a new worksheet. The variable labeled ‘intercept’ will be equal to the natural log of coefficient B. The remaining coefficients (c,d, and e) will be given as X Variable 1,X Variable 2 and X Variable 3 respectively. This process is quick and accurate for Excel 2002 and was used for all equations given in this report. Excel has to capability to handle one dependent variable and multiple independent variables. As many as 12,000 points were used in the regressions.
APPENDIX B

COMPLETE GRI-MECH VERSION 2.1 (REF. 10) CH4 MECHANISM USED

&PRTYPE GLOBAL=.TRUE., GRONLY=.FALSE., &END

2.00 = O2 M 1.200E+17 -1.000 .00

THIRD BODY
H2 O2 2.40 15.40 CH4 2.00 CO 1.75
CO2 C2H6 3.60 3.00 N2 0.83 END
O H = OH M 5.000E+17 -1.000 .00

THIRD BODY
H2 O2 2.00 6.00 CH4 2.00 CO 1.50
CO2 C2H6 2.00 3.00 N2 0.5 END
O H2 = H OH O2 5.000E+04 2.670 6290.00
O HO2 = OH O2 2.000E+13 .000 .00
O H2O2 = OH HO2 9.630E+06 2.000 4000.00
O CH = H CO 5.700E+13 .000 .00
O CH2 = H HCO 8.000E+13 .000 .00
O CH3 = H CH2O 8.430E+13 .000 .00
O CH4 = OH CH3 1.020E+09 1.500 8600.00
O CO = CO2 M 6.020E+14 .000 3000.00

THIRD BODY
H2 O2 6.00 CH4 2.00 CO 1.75
CO2 C2H6 3.00 N2 0.0 END
O H2 = H OH O2 5.000E+13 .000 .00
O HO2 = OH O2 2.000E+13 .000 .00
O CH2O = OH HCO 3.900E+13 .000 3540.00
O CH3O = OH CH2O 1.000E+13 .000 .00
O CH3OH = OH CH2O 1.000E+13 .000 .00
O CH3OH = OH CH2OH 3.900E+13 .000 3540.00
O CH2OH = OH CH2O 3.900E+13 .000 3540.00
O CH3O = OH CH2O 1.000E+13 .000 .00
O C2H = CH CO 5.000E+13 .000 .00
O C2H2 = H HCCO 1.020E+07 2.000 1900.00
O C2H2 = OH C2H 4.600E+19 -1.410 28950.00
O C2H2 = CO CH2 1.020E+07 2.000 1900.00
O C2H3 = H CH2CO 3.000E+13 .000 .00
O C2H4 = CH3 HCO 1.920E+07 1.830 2200.00
O C2H5 = CH3 CH2O 1.320E+14 .000 .00
O C2H6 = OH C2H5 8.980E+07 1.920 5690.00
O HCCO = H 2.0CO 1.000E+14 .000 .00
O CH2CO = OH HCCO 1.000E+13 .000 8000.00
O CH2CO = CH2 CO2 1.750E+12 .000 1350.00
O CO = O CO2 2.500E+12 .000 47800.00
O CH2O = HCO HCO 1.000E+14 .000 40000.00
H O2 = HO2 M 2.800E+18 -.860 .00

THIRD BODY
O2 H2O 1.50 CO .75 CO2 1.50
C2H6 N2 0.0 END
H 2.002 = H2O 2.00 O2 3.000E+20 -1.720 .00
H O2 = O OH 8.300E+13 .000 14413.00
H H = H2 M 1.000E+18 -1.000 .00

THIRD BODY
H2 H2O 2.00 CH4 2.00 CO2 0.63
C2H6 N2 0.0 END
H 2.00 H2 = 2.0H2 9.000E+16 -.600 .00
H 2.0H = H2 H2O 6.000E+19 -1.250 .00
THIRDBODY
H       C2H3    =  C2H4        M          6.080E+12   .270     280.00
THIRDBODY
H2      2.00   H2O       6.00   CH4      2.00   CO      1.50
CO2   2.00   C2H6     3.00   END
H       C2H3    =  H2          C2H2      3.000E+13   .000     .00
H       C2H4    =  C2H5        M          1.080E+12   .454    1820.00
THIRDBODY
H2      2.00   H2O       6.00   CH4      2.00   CO      1.50
CO2   2.00   C2H6     3.00   N2       0.7   END
H       C2H4    =  C2H3        H2      1.325E+06  2.530    12240.00
H       C2H5    =  C2H6        M          5.210E+17  -.990    1580.00
THIRDBODY
H2      2.00   H2O       6.00   CH4      2.00   CO      1.50
CO2   2.00   C2H6     3.00   N2       0.7   END
H       C2H5    =  H2          C2H4      2.000E+12   .000     .00
H       C2H6    =  C2H5        H2      1.150E+08  1.900    7530.00
H       CH2CO   =  HCCO        H2       5.000E+13   .000    8000.00
H       CH2CO   =  CH3        CO      1.130E+13   .000    3428.00
H       HCCOH   =  H          CH2CO     1.000E+13   .000     .00
H2      CO      =  CH2O        M          4.300E+07  1.500    79600.00
THIRDBODY
H2      2.00   H2O       6.00   CH4      2.00   CO      1.50
CO2   2.00   C2H6     3.00   N2       0.7   END
OH      H2      =  H          H2O       2.160E+08  1.510    3430.00
OH      OH      =  H2O2       M          7.400E+13  -.370     .00
THIRDBODY
H2      2.00   H2O       6.00   CH4      2.00   CO      1.50
CO2   2.00   C2H6     3.00   N2       0.7   END
2.0OH      =  O          H2O       3.570E+04  2.400    -2110.00
OH      H2O2    =  H2O2       H2O      1.750E+12   .000    320.00
OH      H2O2    =  H2O2       H2O      5.800E+14   .000    9560.00
OH      C       =  H          CO       5.000E+13   .000     .00
OH      CH      =  H          HCO      3.000E+13   .000     .00
OH      CH2     =  CH2        H2O      2.000E+13   .000     .00
OH      CH2     =  CH2        H2O      1.130E+07  2.000    3000.00
OH      CH3     =  CH3OH      M          6.300E+13   .000     .00
THIRDBODY
H2      2.00   H2O       6.00   CH4      2.00   CO      1.50
CO2   2.00   C2H6     3.00   N2       0.7   END
2.0OH      =  O          H2O       3.570E+04  2.400    -2110.00
OH      H2O2    =  H2O2       H2O      1.750E+12   .000    320.00
OH      H2O2    =  H2O2       H2O      5.800E+14   .000    9560.00
OH      C       =  H          CO       5.000E+13   .000     .00
OH      CH      =  H          HCO      3.000E+13   .000     .00
OH      CH2     =  CH2        H2O      2.000E+13   .000     .00
OH      CH2     =  CH2        H2O      1.130E+07  2.000    3000.00
OH      CH3     =  CH3OH      M          6.300E+13   .000     .00
THIRDBODY
H2      2.00   H2O       6.00   CH4      2.00   CO      1.50
CO2   2.00   C2H6     3.00   N2       0.7   END
OH      CH3     =  CH2        H2O       5.600E+07  1.600    5420.00
OH      CH4     =  CH3        H2O       1.000E+08  1.600    3120.00
OH      CO      =  H          CO2      4.760E+07  1.228     70.00
OH      HCO     =  H2O        CO       5.000E+13   .000     .00
OH      CH2O    =  HCO        H2O      3.430E+09  1.180    -447.00
OH      CH3O    =  H2O        CH2O     5.000E+12   .000     .00
OH      C2H     =  H          HCCO      2.000E+13   .000     .00
OH      C2H2    =  H          CH2CO     2.180E-04  4.500    -1000.00
OH      C2H2    =  C2H        H2O      3.370E+07  2.000    14000.00
OH      C2H2    =  CH3        CO       4.830E-04  4.000    -2000.00
OH      C2H3    =  H2O        C2H2     5.000E+12   .000     .00
OH      C2H4    =  C2H3       H2O      3.600E+06  2.000    2500.00
OH      C2H6    =  C2H5       H2O      3.540E+06  2.120    870.00
OH      CH2CO   =  HCCO       H2O      7.500E+12   .000    2000.00
2.0HO2     =  O2        H2O2       1.300E+11   .000    -1630.00
2.0HO2     =  O2        H2O2       4.200E+14   .000    12000.00
H2O      CH2     =  OH        CH2O      2.000E+13   .000     .00
HO2 + CH3 = O2 + CH4  1.000E+12   .000   .00
HO2 + CH3 = OH + CH3O  2.000E+13   .000   .00
HO2 + CO = OH + CO2  1.500E+14   .000   23600.00
HO2 + CH2O = OH + CH2O   2.000E+13   .000   .00
C + O2 = O + CO   5.800E+13   .000   576.00
C + CH2 = H + C2H   5.000E+13   .000   .00
C + CH3 = H + C2H2  5.000E+13   .000   .00
CH + O2 = O + HCO   3.300E+13   .000   .00
CH + H2 = H + CH2   1.107E+08   1.790   1670.00
CH + H2O = H + CH2O  1.713E+13   .000  -755.00
CH + CH2 = H + C2H2  4.000E+13   .000   .00
CH + CH3 = H + C2H3  3.000E+13   .000   .00
CH + CH4 = H + C2H4  6.000E+13   .000   .00
CH + CO = HCCO + M   5.000E+13   .000   .00

THIRDBODY

H2  2.00  H2O  6.00  CH4  2.00  CO  1.50
CO2  2.00  C2H6  3.00  N2  0.7  END
CH  CO2 = HCO + CO   3.400E+12   .000   690.00
CH  CH2O = H + CH2CO  9.460E+13   .000  -515.00
CH  HCCO = CO + C2H2   5.000E+13   .000   .00
CH  CH2 = OH + HCO   1.320E+13   .000   1500.00
CH  H2 = H + CH3   5.000E+05   2.000   7230.00
2.0CH2 = H2 + C2H2   3.200E+13   .000   .00
CH2  CH3 = H + CH2O  4.000E+13   .000   .00
CH2  CH4 = 2.0CH3  2.460E+06   2.000  8270.00
CH2  CO = CH2CO + M   8.100E+11   .500   4510.00

THIRDBODY

H2  2.00  H2O  6.00  CH4  2.00  CO  1.50
CO2  2.00  C2H6  3.00  N2  0.7  END
CH  HCCO = C2H3 + CO   3.000E+13   .000   .00
CH  O2 = O + CH3O  2.675E+13   .000  28800.00
CH  O2 = OH + CH2O  3.600E+10   .000   8940.00
CH  H2O2 = HO2 + CH4   2.450E+04   2.810   5860.00
2.0CH3 = C2H6 + H2   3.200E+13   .000   .00

THIRDBODY

H2  2.00  H2O  6.00  CH4  2.00  CO  1.50
CO2  2.00  C2H6  3.00  N2  0.7  END
2.0CH3 = H + C2H5  4.990E+12   .100  10600.00
CH  HCO = CH4 + CO   2.648E+13   .000   .00
CH  CH2O = HCO + CH4   3.320E+03   2.810   5860.00
CH  CH3OH = CH2OH + CH4   3.000E+07   1.500   9940.00
CH  CH3OH = CH3O + CH4   1.000E+07   1.500   9940.00
CH  C2H4 = C2H3 + CH4   2.270E+05   2.000   9200.00
CH  C2H6 = C2H5 + CH4   6.140E+06   1.740  10450.00
HCO  CH2OH  7.600E+12   .000   400.00
CH2OH  CH2O  1.800E+13   .000   900.00
CH3O  CH3OH  4.280E+13   7.600  -3530.00
C2H  O2 = H + CH2H  4.070E+05   2.400   200.00
C2H3  O2 = HCO + CH2O   3.980E+12   .000  -240.00
M = C2H4 = H + C2H2  8.000E+12   .440   88770.00

THIRDBODY

H2  2.00  H2O  6.00  CH4  2.00  CO  1.50
CO2  2.00  C2H6  3.00  N2  0.7  END
C2H5  O2 = HO2 + C2H4  8.400E+11   .000   3875.00
HCCO  O2 = OH + 2.0CO  1.600E+12   .000   854.00

NASA/TM—2004-213046  26
2.0HCCO =2.0CO C2H2 1.000E+13 .000
330.00
N NO = N2 O 3.500E+13 .000
6400.00
N O2 = NO O 2.650E+12 .000
N2O OH = NO2 H 7.333E+13 .000 1120.00
N2O O = NO2 2.900E+13 .000 23150.00
N2O H = N2 OH 4.400E+14 .000 18880.00
N2O OH = NO2 H2 2.000E+12 .000 21060.00
M N2O = N2 O 1.300E+11 .000 59620.00
THIRDBODY
H2 2.00 H2O 6.00 CH4 2.00 CO 1.50
CO2 2.00 C2H6 3.00 N2 0.625 END
HO2 NO = NO2 OH 2.110E+12 .000 -480.00
NO O = NO2 M 1.060E+20 -1.410 .00
THIRDBODY
H2 2.00 H2O 6.00 CH4 2.00 CO 1.50
CO2 2.00 C2H6 3.00 END
NO2 O = NO O2 3.900E+12 .000 -240.00
NO2 H = NO OH 1.320E+14 .000 360.00
NH O = NO H 5.000E+13 .000 .00
NH H = N H2 3.200E+13 .000 330.00
NH OH = HNO H 2.000E+13 .000 .00
NH OH = N H2O 2.000E+09 1.200 .00
NH O2 = HNO O 4.610E+05 2.000 6500.00
NH O2 = NO OH 1.280E+06 1.500 100.00
NH N = N2 H 1.500E+13 .000 .00
NH H2O = HNO H2 2.000E+13 .000 13850.00
NH NO = N2 OH 2.160E+13 -.230 .00
NH NO = N2O H 4.160E+14 -.450 .00
NH2 O = OH NH 7.000E+12 .000 .00
NH2 O = H HNO 4.600E+13 .000 .00
NH2 H = NH H2 4.000E+13 .000 3650.00
NH2 OH = NH H2O 9.000E+07 1.500 -460.00
M NNH = N2 H 3.300E+08 .000 .00
M NNH = N2 H 1.300E+14 -.110 4980.00
THIRDBODY
H2 2.00 H2O 6.00 CH4 2.00 CO 1.50
CO2 2.00 C2H6 3.00 END
NNH O2 = HO2 N2 5.000E+12 .000 .00
NNH O = OH N2 2.500E+13 .000 .00
NNH O = NH NO 7.000E+13 .000 .00
NNH H = H2 N2 5.000E+13 .000 .00
NNH OH = H2O N2 2.000E+13 .000 .00
NNH CH3 = CH4 N2 2.500E+13 .000 .00
H NO = HNO M 8.950E+19 -1.320 740.00
THIRDBODY
H2 2.00 H2O 6.00 CH4 2.00 CO 1.50
CO2 2.00 C2H6 3.00 END
HNO O = NO OH 2.500E+13 .000 .00
HNO H = H2 NO 4.500E+11 .720 660.00
HNO OH = NO H2O 1.300E+07 1.900 -950.00
HNO O2 = HO2 NO 1.000E+13 .000 13000.00
CN O = CO N 7.700E+13 .000 .00
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<th>Species 2</th>
<th>Product</th>
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<th>Energy 2</th>
<th>Energy 3</th>
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<td>CO$_2$ + 2.00 C$_2$H$_6$ + 3.00 N$_2$</td>
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<td>CO$_2$ + 2.00 C$_2$H$_6$ + 3.00 N$_2$</td>
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<td>N2</td>
<td>+ OH</td>
<td>2.00E+15</td>
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</table>

H + O2 = H2O = H2O

1. H O2 N2 = H2O N2
2. OH CH2 OH = H2O CH2O
3. OH CH3 OH = CH2 OH H2O
4. OH CH3 OH = CH3O H2O
5. HCO H2O = H CO H2O
6. NCO OH = NO H CO
7. NCO OH = NO H CO
MODIFICATIONS TO GLSENS FOR WATER INJECTION

C WATER INJECTION MODIFICATION FOR CH4 TO SUBROUTINE INIT
15 IF (FUEL.EQ.BLANK) GO TO 80
C SPECIAL COMPOSITION INPUT
   DO 20 J=1,NRS
      IF (FUEL.NE.DSPNM(J)) GO TO 20
      JFUEL=J
      GO TO 30
20 CONTINUE
   WRITE (LWRITE,25) FUEL
25 FORMAT(/,7H0(INIT),5X,'***** WARNING: INPUT FUEL NAME "',A8,
1 '" IS NOT IN REACTANT LIST *****',//)
   NEXT=.TRUE.
   RETURN
30 IF (ERATIO.NE.0.) GO TO 35
   IF (FLAIR.EQ.0.) GO TO 80
C FUEL-OXIDANT MASS RATIO SPECIFIED
   TERM1 = AIRMW*FLAIR/MW(JFUEL)
   FRO2 = 1.0/(1.0 + NOXRAT + ARAT + CRAT)
   TERM2 = 1.0/(1.0 + TERM1)
   CONCO2 = DBLE(FRO2*TERM2)
   GO TO 40
C FUEL-OXIDANT EQUIVALENCE RATIO SPECIFIED
35 H2OMF=H2OF*16./18.
   TERM1 = 4.0*ERATIO*(1.+H2OMF)
   TINP = 4.0*SCC + SCH - 2.0*SCOX
   ERAT=ERATIO
   TERM2 = 1.0 + NOXRAT + ARAT + CRAT
   FRO2 = 1.0/TERM2
   TERM2 = 1.0/(TERM1 + TINP*TERM2)
   CONCO2 = DBLE(TINP*TERM2)
C
40 IF (MOLEF) GO TO 42
C WARNING: MOLEF = .FALSE. FOR SIMPLIFIED COMPOSITION INPUT.
   SET MOLEF = .TRUE.
   MOLEF = .TRUE.
   WRITE (LWRITE,41)
41 FORMAT(/,7H0(INIT),5X,'**** WARNING: SIMPLIFIED COMPOSITION ',
1 'INPUT USED BUT MOLEF SET EQUAL TO .FALSE. IN NAMELIST START',//,
2 30X,'MOLEF HAS BEEN SET EQUAL TO .TRUE. ****',//)
C SIGMA = MOLES/GRAM MIXTURE
42 SIGMA(JFUEL) = DBLE((TERM1*TERM2)/(dble(1.+H2OMF)))
   DO 45 I = 1,LS
      IF (O2.NE.DSPNM(I)) GO TO 45
      SIGMA(I) = CONCO2
      GO TO 50
45 CONTINUE
50 SIGMA(27) =SIGMA(JFUEL)*DBLE(H2OMF)
   DO 55 I = 1,LS
      IF (N2.NE.DSPNM(I)) GO TO 55
      SIGMA(I) = DBLE(NOXRAT)*CONCO2
      GO TO 60
55 CONTINUE
Additional equations specific to the program.

\[
Term2 = \frac{1}{Term1 + TINP + Term2}
\]  \hspace{1cm} (C1)

\[
\text{conc } O_2 = y_{O_2} = \frac{TINP * 1}{Term2}
\]  \hspace{1cm} (C2)

\[
y_f = \frac{4 \text{Eratio}}{4n_c + n_H - 2n_O} \times y_{O_2} = \frac{Term1}{TINP} \times TINP \times Term2
\]  \hspace{1cm} (C3)

\[
y_{H2O} = y_{fuel} \times H2OF \times \frac{167}{18}
\]  \hspace{1cm} (C4)

\[
y_{N_2} = \eta_{N_2} \times y_{O_2}
\]  \hspace{1cm} (C5)

let \( y_{Ar} = 0.0 \), and \( y_{CO2} = 0.0 \)
APPENDIX D
MODIFICATIONS TO OUT2 OF GLSENS FOR OBTAINING OUTPUT FOR THE CH4 COMPUTATIONS

C ALREADY DONE IN DIFFUN
C ch4.f dec 14,2003
420 foa=eratio*0.059
tn=timil
trr=sngl(T)
418 if(time.gt.0.)goto 427
write(10,1)
1 format(’ TITLE= P atm T K f/a conc fuel co nox coequil time ’)
C calculate the initial conditions for the averaging
write(11,417)ch4e
417 format(’ch4e ’,e13.4)
nc=1
nco=1
t0=0.
stco=0.
areaf=0.
tauf=0.
tauco=0.
tauno=0.
areaco=0.
areano=0.
atauf=0.
atauco=0.
atauno=0.
427 continue
cco=sngl(dabs(prc(9)))
cnox=sngl(dabs(prc(31)+2.d0*prc(33)+prc(34)))
xnox=sngl(prc(31)/(prc(31)+2.d0*prc(33)+prc(34)))
cfuel=sngl(dabs(prc(14))))
C write(10,423)nc,P,T,foa,cfuel,cco,cnox,ccoe,
C 1timil
423 format(i4,f5.1,f8.1,1p,8e10.3)
C Begin Step One calculations
C calculate the initial conditions for the averaging
if(nc.lt.2)goto 1500
if(nc.eq.2)tstart=timil
if(nc.eq.2)t0=timil
O2=sngl(prc(2))
if(nc.eq.2)tauf=-sngl(prc(14)/W(14))*1.e3
if(W(14).ne.0.)tauf=-sngl(prc(14)/W(14))*1.e3
if(tauf.le.0.)goto 424
C if(cfuel.lt.1.e-25)goto 424
if(W(14).eq.0.)tauf=-sngl(prc(14)/W(14))*1.e3
if(tauf.le.0.)goto 424
C if(nc/10*10.eq.nc)write(11,423)nc,P,T,foa,cfuel,cco,cnox,ccoe,
C 1timil
424 format(i4,f5.1,f8.1,1p,8e10.3)
if(nc.eq.2) tauni=sngl(1.D0/(W(31)+2.d0*W(33)+W(34)))*1.e3
   tauno=sngl(1.D0/(W(31)+2.d0*W(33)+W(34)))*1.e3
   if(t0.eq.tn) goto 339
   if(tauno.lt.0.) go to 339
   areano=areano+(1./tauno+1./tauni)/2.*(tn-t0)
   atauno=timet/areano
   tauni=tauno
   xrate=sngl(W(31)+2.d0*W(33)+W(34))
   if((nc/10*10).eq.nc)write(13,501)nc,P,T,xrate,cnox,tauno,xnox,
   1atauno,eratio,timil
339 denm=sngl(W(15))
   cco2=sngl(prc(15))
   C if(w(15).lt.0.)next=.true.
   if(nc.eq.2) tauci=(sngl(dabs(prc(9))))/denm*1.e3
   tauco=(sngl(dabs(prc(9))))/denm*1.e3
   if(t0.eq.tn) goto 504
   if(tauco.le.0.) goto 503
   C if(stco.eq.0.)stco=timil
   timco=timil
   rfuel=cfuel/tauf*1.e3
   areaco=areaco+(1./tauco+1./tauci)/2.*(tn-t0)
   atauc=timco/areaco
   tauco=tauci
   nco=nco+1
338 format(i7,1p8e11.3)
   rfuel2=sngl(w(14))
   dcodt=-(sngl(prc(9)))-cco*1.e3
503 if((nc/10*10).eq.nc)write(12,423) nc,P,T,foa,cco,ccoe,tauco,
   1atauco,timil
if(nco.lt.100)write(15,338)nc,timco,tauci,tauco,atauco,areaco
504 t0=tn
431 format (f12.2,3e13.5,f8.3,f8.3,e13.5)
501 format(i4,f7.4,f7.1,4e12.5,f6.3,f9.3)
C Begin step Two calculations
   pl=log(P)
   tol=1./T
   h2ol=prc(27)
   h2osl=sngl(prc(27))
   h2ol=dlog(prc(27))
   C if(cfuel.lt.1.e-20)cfuel=1.e-20
   fo=log(cfuel)
   fo2=sngl(prc(2))
   fo2l=log(fo2)
   tal=log(tauf)
   talc=log(tauci)
   taln=log(tauni)
   if(nc.eq.2) tolo=log(tfop)
   if(nc.gt.2) tolo=timil
   if(cfuel.le.1.e-24) goto 902
425 format(i6,1p9e11.3)
C output fuel times
   write(26,900)foa,h2of,pl,tol,fo,fo2l,h2ol,tal,tolo
900 format(2f7.3,1p,9e11.3)
902 tal=log(tauci)
   if(cco.lt.1.e-20) goto 903
   col=log(cco)
   tolo=log(tcop)
if(nc.gt.2) tolo=timil
C Output CO times and concentrations
   write(27,900)foa,h2of,pl,tol,fo,col,fo2l,h2ol,tal,tolo
903   continue
   tal=log(tauni)
   tolo=log(tnop)
   if(nc.gt.2) tolo=timil
C output NOx times and concentrations
   write(28,900)foa,h2of,pl,tol,fo,fo2l,h2ol,tal,tolo
1500  nc=nc+1
C   if(nc.gt.4000) next=.true.
C   if(nc.gt.50) next=.true.
   if (timil.gt.1.) go to 502
DO 435 IJ=1,MAX
   IF (IJ.GT.LS.OR.IJ.GT.LR) GO TO 435
   TCON(IJ)=SNGL(PRC(IJ)/W(IJ))
   FMOL=SNGL(SIGMA(IJ)*MIXMW)
   WRITE (LWRITE,175) DSPNM(IJ),PRC(IJ),FMOL,W(IJ)
GO TO 435
430   WRITE (LWRITE,185) IJ,RATE(IJ),PRX(IJ),EQUIL(IJ)
C430   continue
435   CONTINUE
APPENDIX E

RUN INPUT FOR GLSENS

Run input for Jet-A for GLSENS TYPICAL FOR STEP ONE
TIME        PRESSURE      C12H23
&prob TCON=.TRUE.,CT0=1.0,
IPRINT=1,end=1.e-2,,&end
&start time=0.,ERATIO=0.3,SCC=12.0,SCOX=0.,SCH=23.,
NOXRAT=3.76,ARAT=0.,CRAT=0.,T=1000.,H2OF=0.&end
END
&solver mxstep=2000.,emax=1.e-13,atolsp=1.e-13,maxstp=190000,&end
INIT
ALLSP     END
SENSVAR
ALLSP     END
REAC
FINIS

RUN INPUT FOR JET-A FOR GLSENS TYPICAL FOR STEP TWO
REPEAT  MECHANISM
TIME        PRESSURE      C12H23
&prob TCON=.TRUE.,CT0=1.0,
PRINT=5.e-5,1.e-4,1.5e-4,2.e-4,2.5e-4,
3.e-4,4.e-4,5.e-4,6.e-4,7.e-4,8.e-4,9.e-4,1.e-3,1.5e-3,2.e-3,2.5e-3,3.e-3,3.5e-3,4.e-3,
4.5e-3,5.e-3,5.5e-3,6.0e-3,&end
&start time=0.,ERATIO=0.3,SCC=12.0,SCOX=0.,SCH=23.,
NOXRAT=3.76,ARAT=0.,CRAT=0.,T=1000.,H2OF=0.&end
END
&solver mxstep=2000.,emax=1.e-13,atolsp=1.e-13,maxstp=190000,&end
INIT
ALLSP     END
SENSVAR
ALLSP     END
REAC
FINIS

RUN INPUT FOR CH4 FOR GLSENS TYPICAL FOR STEP ONE
TIME        PRESSURE      CH4
&prob TCON=.TRUE.,CT0=1.0,
IPRINT=1,end=5.e-4,,&end
&start time=0.,ERATIO=0.4,SCC=1.0,SCOX=0.,SCH=4.,
NOXRAT=3.76,ARAT=0.,CRAT=0.,T=1000.,H2OF=0.&end
END
INIT
ALLSP     END
SENSVAR
ALLSP     END
REAC
FINIS

RUN INPUT FOR CH4 FOR GLSENS TYPICAL FOR STEP TWO
TIME        PRESSURE      CH4
&prob TCON=.TRUE.,CT0=1.0,
PRINT=5.e-6,1.e-5,2.5e-5,5.e-5,1.e-4,1.5e-4,2.e-4,2.5e-4,
3.e-4,4.e-4,5.e-4,1.e-3,1.5e-3,2.e-3,2.5e-3,3.e-3,3.5e-3,4.e-3, 
4.5e-3,5.e-3,5.5e-3,6.0e-3,&end

&start time=0.,ERATIO=0.3,SCC=1.0,SCOX=0.,SCH=4.0, 
NOXRAT=3.76,ARAT=0.,CRAT=0.,T=1000.,H2OF=0.0,&end
END
&solver mxstep=20000,emax=1.e-14,atolsp=1.e-14,maxstp=300000, &end
ALLSP     END
SENSVAR
ALLSP     END
REAC
FINIS
APPENDIX F

FORTRAN CODE FOR TANKS IN SERIES CALCULATION

C IN FIGURE 8 OUTPUT THE PPM AS A FUNCTION OF TEMPERATURE AND COMPUTE
C EACH PHIN TO KEEP TEMPERATURE CONSTANT AS H INCREASES
C methane
C H=h/a t=time
C set up maximum reactors to be eight in series
C pf=f/a fraction in reactor, t=time milliseconds
DIMENSION pf(8),t(8),TT(8),EFT(8),TAUF(8),TCO(8),TNOX(8),tim(8)
DIMENSION fa(8),COE(8),rnox(10),hn(5),melissa(5)
data pf/1.5,1.2,1.0,5*0./
data t/1.,1.5,2.5,5*0.5/
REAL NOX,melissa
open(5,file='fort.5',form='formatted')
open(10,file='fort.10',form='formatted')
open(11,file='fort.11',form='formatted')
open(12,file='fort.12',form='formatted')
open(13,file='fort.13',form='formatted')
open(14,file='fort.14',form='formatted')
open(15,file='fort.15',form='formatted')
write(12,111)
C 111 format(' h2o  phi  h2o/f  eftt  tt(4)  ppmco  ',
C     ' phi ',8X, 'h2o', 2X 'T4 ',9x,'O2    O2COR ',
C     ' ppmno,  cwcd')
T3=700.
P=200./14.696
C   read(5,5)p3,T3,pf,t
5 format(2e10.3,/,8f5.2,/8f5.3)
c p=p3/14.696
WRITE(10,7)P,T3,((PF(I),T(I)),I=1,4)
WRITE(13,7)P,T3,((PF(I),T(I)),I=1,4)
7 FORMAT(' P=',F5.2,' T3= ',F5.0,/,' PF,T=',2(2F5.2' : ')/2(2F5.2' : ')
C     1'),/3x,'TEMP F H=0.0 H=0.05 H=0.1 H=0.15 ')
WRITE(15,7)P,T3,((PF(I),T(I)),I=1,4)
WRITE(10,8)
8 FORMAT(6x,'IR   h   FA(IR)  T4(IR)   EFF     nox
C     1 Ttnox TAUNOX')
PHI=0.2
DO 11 IP=1,8
PHI=PHI+0.1
H=-0.05
DO 1 IH=1,4
H=H+0.05
PHIN=PHI
DO 36 IX=1,10
36 PHIN=PHIN*1.4/(PHIN*0.059)**0.1586
C WRITE(10,37)IX,PHIN,PHI,H
37 FORMAT(I5,3F11.4)
pf(1)=1./phin
if(phin.ge.1.)write(13,17) Phi, Phin,H
17 format(' phi=',e12.4,'phim=',e12.4,' h',e12.4)
if(PHIN.GT.1.) go to 500
CO=0.
NOX=0.
C ITERATE TO SOLVE FOR T4 FOR EACH REACTOR, IR= CONDITIONS FOR EACH REACTOR.
(NO)

tim(1)=0.
DO 30 IR=1,3
   tim(ir)=tim(ir)+t(ir)
   FA(ir)=PHIN*0.059*PF(IR)
   IF(fa(ir).gt.0.059)fa(ir)=0.059
   HF=H/FA(ir)
   IF(HF.GT.2.5)HF=2.5
   EFF=100.00
   EFTT=100.
   itI=0
C DO 40 IT4=1,10
   itI=itI+1
C EFTT= TOTAL INEFFICIENCY
   EFFO=EFF
C DO NOT USE RICH CORRELATIONS HERE AS RESULTS FLUCTUATE – NOT PHYSICAL
   IF(fa(ir).GT.0.0592) GO TO 66
C DO LEAN CORRELATIONS
   T4E=457.*(T3)**0.258*FA(ir)**0.435*(1.+HF)**(-.069)*(P)
1**0.00284*0.059**(-.435)
   eff=100.
   T4=T3+(T4E-T3)*EFF/100.
   TAU(IF)=2.09e-4*FA(ir)**0.222*(1.+HF)**0.0675*P**(-1.07)*
   1EXP(22625./T4)
   TCO(IR)=9.99E-4*FA(ir)**(0.101)*(1.+HF)**.0959*p**(-1.00)*
   1EXP(2434./T4)
   TNOX(IR)=2.94E4*fa(ir)**.032*(1.+HF)**.329*p**(-2.1)*
   1EXP(34859./T4)
   COE(IR)=15569.*fa(ir)**1.90*(1.+HF)**.00705*P**(-.477)*
   1EXP(-33388./T4)*0.059**1.90
   goto 40
C DO RICH CORRELATIONS
   66 T4E=859.*T3**0.176*fa(ir)**(-0.559)*(1.+HF)**(-0.146)*P**
   1(.00167)*0.059**.559
   T4=T3+(T4E-T3)*EFF/100.
   TAU(IF)=2.74e-2*fa(ir)**.111*(1.+HF)**(-.829)*P**(-0.328)*
   1EXP(14216./T4)
   TCO(IR)=1.3E-3*fa(ir)**.215*(1.+HF)**.195*p**(-1.0)*
   1EXP(32649./T4)
   TNOX(IR)=43928.*fa(ir)**.284*(1.+HF)**.641*p**(-2.0)*
   1EXP(-31365./T4)*0.059**(-2.55)
C EFF EQUALS THE EFFICIENCY OF EACH REACTOR
C IF CHANGE IN EFFICIENCY IS LESS THAN 0.1 ITS CONVERGED
C 67 EFF=(1.-EXP(-tim(IR)/TAUF(IR)))*100.
C IF(ABS(EFP-EFFO).LT.0.01)GO TO 41
C WRITE(10,96)IR,iTI,fa(ir),T4,EFF
40 CONTINUE
   96 FORMAT(' IR= ',I10,' itI ',i4,' fa ',e10.3,' T4 ',F7.2,F7.2)
41 TT(IR)=T4
   EFT(IR)=EFF
   FAE=fa(ir)*EFF/100.
   XFO=FAE/(FAE*(1.+HF*16./29.))+16./29.
   FO=XFO*P/82.056/T4
   eftt=eftt*(100.-eff)
C write(10,97) ir,phi,hf
         97 format(' ir ',i5,' phi ',f10.3,' hf ',f8.3)
100 FORMAT(2X,4e12.4)
c CALCULATE CO FOR EACH REACTOR
TTF=T(IR)/TAUF(IR)
TTCO=T(IR)/TCO(IR)
EXCO=EXP(-TTCO)
TTAUF=T(IR)/TAUF(IR)
        FOT=FO*TTAUF/(TTCO-TTAUF)
        if(ir.eq.1) CO=COE(IR) - EXCO*COE(IR) + FOT*(EXP(-TTAUF) - EXCO)
        irb=ir
        IF(IR.GT.1) irb=ir-1
CO=CO*FA(IR)/FA(IRB)*(1.+fa(irb)*29./16.)/(1.+fa(ir)*29./16.)
        IF(IR.GT.1) CO=COE(IR) + (CO-COE(IR))*EXCO + FOT*(EXP(-TTAUF) - EXCO)
C CALCULATE NOX FOR EACH REACTOR
TTNOX=T(IR)/TNOX(IR)
        NOX=NOX*FA(IR)/FA(IRB)*(1.+fa(irb)*29./16.)/(1.+fa(ir)*29./16.)
        IF(IR.GT.1) NOX=NOX + TTnox
phio=fa(ir)/0.059
WRITE(10,99)IR,IP,H,phio,T4e,EFF,NOX,TTNOX,TNOX(IR)
99 FORMAT(*',2I4,2F8.4,2F9.2,1P6E10.2)
110 rnox(ih)=nox*1.e6*82.056*t4/p
write(11,98)ir,PHIO,t(ir),ttf,ttco,CO,COE(IR),T4
98 format(i3,1p7e10.2)
c END OF INDIVIDUAL REACTOR SEGMENT
30 continue
113 format(2x,f9.3,1p5e11.3)
CWCD=(PHIN+(2.+(16./18.)*(HF)*PHIN)+(1.-PHIN)*2+(2.*79./21.))/
     (1.(PHIN+(1.-PHIN)*2+(2.*79./21.))
RHO=P/82.056/TT(3)
O2=(1.-PHIN)*2.*phin/(9.524-phin)
if(O2.lt.0.)O2=0.
O2COR=5.9/(20.9-100.*O2)
PPMCO=CO*1.E6/RHO*O2COR
PPMNO=NOX*1.E6/RHO*O2COR
hnh=ppmno
melissa(ih)=ppmco
write(12,102)phi,hf,tt(3),o2,o2cor,ppmno,cwcd
102 format(2f8.3,f7.1,f8.3,1p3e10.2)
c write(13,113)phi,(rnox(ir),ir=1,4)
1 continue
Tf=tt(3)*1.8-460.
write(13,113)tf,(hn(in),in=1,4)
write(14,113)tf,(rnox(in),in=1,4)
write(15,113)tf,(melissa(in),in=1,4)
11 continue
500 stop
End
APPENDIX G

FORTRAN CODE FOR MODEL SIMULATION OF GLSENS CONSTANT TEMPERATURE AND PRESSURE PLUG FLOW REACTOR

C Moder Ch4 reaction
OPEN(6,FILE='MELISSA.OUT',FORM='FORMATTED')
foa=0.05263
TK=2000.
P=2.
  rho=P/82.056/TK
  yfo=foa/(foa+16./29.)
  fo=yfo*rho
C Step one equation - lean
  tauf1=2.09e-4*P**(-1.07)*foa**(0.222)*exp(22625./TK)
  TMIN=-TAUF1*LOG(1.-1.E-20/2./FO)
WRITE(6,1)TMIN,TAUF1,fo
1 FORMAT(' TMIN MS ',E12.4,' TAUF1 MS ',E12.4,' fo ',e12.4)
write(6,2)
2 format(8X,' TIME SEC  FUEL    H2O     TAUF MS ')
C CH4+2.O2=CO2+2.H2O
NT=0
F=FO
H2O=0.
TIME=0.
TAUF=TAUF1
DTIME=1.E-5
DO 6 IT=1,50000
  TIME=TIME+DTIME
  DF=F*DTIME/TAUF
  F=F-DF
  H2O=H2O+2*DF
C step two equation
  IF(H2O.GT.1.E-20)TAUF=3.5E-9*P**(-0.0713)*F**(-0.307)*
  2H2O**(-0.465)*(1.+H2O/F)**(-0.221)*EXP(14149/TK)
  TSEC=TIME*1.E-3
  NT=NT+1
  IF (IT/100*100.EQ.IT)WRITE(6,10)NT,TSEC,F,H2O,TAUF
10 FORMAT(I8,1P4E12.4)
6 CONTINUE
STOP
END
REFERENCES


Figure 1
Magnussen Mixing Model

Chemical Kinetic Rate

Mixing Rate

Transition Temperature

Rate

Mixing Rate

Temperature
Figure 2
Equilibrium Jet A Water Injection COeq Parity
0.2 < Eratio ≤ 1.0
(lean)
Figure 3
Kinetic Jet-A Water Injection Fuel Tau Parity
Step One
(lean)

NASA/TM—2004-213046
Figure 4
Kinetic Jet-A Water Injection CO Tau Parity
Step One
(lean)
Figure 5
Kinetic Jet-A Water Injection Nox Tau Parity
Step One
(lean)
Figure 6
Jet-A Fuel Step One Lean Model

\[ f/a^{0.238}P^{-0.60}\exp(14202/T) \]

\[ \text{h/f} = 0 \]
\[ \text{h/f} = 0.5 \]
\[ \text{h/f} = 1.0 \]
\[ \text{h/f} = 1.5 \]
\[ \text{h/f} = 2.0 \]
\[ \text{No Water Injection} \]
Figure 7
Kinetic Jet-A Fuel Tau Parity
Step Two
(lean)

![Plot showing tau fuel GLSENS (experimental) vs. tau fuel (predicted). The data points are scattered, with a linear trend line indicating a positive correlation.](image-url)
Figure 9
Kinetic Jet-A NOx Tau Parity
Step Two
(lean)

Tau NOx GLSENS (experimental)

Tau NOx (predicted)
Figure 10
Equilibrium Methane CO Parity
(rich)

0 0.02 0.04 0.06 0.08 0.1 0.12 0.14 0.16
CO eq experimental
CO eq predicted

NASA/TM—2004-213046
Figure 11
Kinetic Methane Fuel Tau Parity
Step One
(lean)
Figure 12
Kinetic Methane CO Tau Parity
Step One
(lean)
Figure 13
Kinetic Methane NOx Tau Parity
Step One
(lean)
Figure 14
Methane Fuel Simple Step One Lean Model

\[(f/a)^{0.222(P)^{-1.07}} \exp(22625/T)\]

- \(h/f=0\)
- \(h/f=0.5\)
- \(h/f=1.0\)
- \(h/f=1.5\)
- \(h/f=2.0\)
Figure 15
Kinetic Methane Tau Fuel Parity
Step Two (lean)

Tau CH4 GLSENS (experimental)

Tau CH4 (predicted)
Figure 16
Kinetic Methane CO Tau Parity
Step Two
(lean)
Figure 17
Kinetic Methane NOx Tau Parity
Step Two
(lean)
Figure 18
NOx water/NOx no water for Jet-A at f/a=0.06

Data
NOxw/NOx
T corrected

fraction weight H2O/Air

NOx/NOx no water

0 0.02 0.04 0.06 0.08 0.1 0.12 0.14 0.16
Figure 19

Methane Data Comparison (f/a=0.055)

- Nox water/Nox no water
- Time Corrected
- T Constant
- Data
Figure 21

Equivalence Ratio vs. CO (ppm) for different values of H:
- H=0.0 (prediction)
- H=0.05 (prediction)
- H=0.10 (prediction)
- H=0.15 (prediction)
- H=0.0 (data)
- H=0.05 (data)
- H=0.10 (data)
- H=0.15 (data)
Figure 24

tauf millisec for lean ch4
Figure 25
CH4 T=2000 P=2 atm Phi=0.8915 Gri-mech viet

CH4 model
CH4 glsens

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Cleveland, Ohio 44135–3191

A simplified kinetic scheme for Jet-A, and methane fuels with water injection was developed to be used in numerical combustion codes, such as the National Combustor Code (NCC) or even simple FORTRAN codes that are being developed at Glenn. The two time step method is either an initial time averaged value (step one) or an instantaneous value (step two). The switch is based on the water concentration in moles/cc of $10^{-20}$ /H$_{11003}$. The results presented here results in a correlation that gives the chemical kinetic time as two separate functions. This two step method is used as opposed to a one step time averaged method previously developed to determine the chemical kinetic time with increased accuracy. The first time averaged step is used at the initial times for smaller water concentrations. This gives the average chemical kinetic time as a function of initial overall fuel air ratio, initial water to fuel mass ratio, temperature, and pressure. The second instantaneous step, to be used with higher water concentrations, gives the chemical kinetic time as a function of instantaneous fuel and water mole concentration, pressure and temperature ($T_4$). The simple correlations would then be compared to the turbulent mixing times to determine the limiting properties of the reaction. The NASA Glenn GLSENS kinetics code calculates the reaction rates and rate constants for each species in a kinetic scheme for finite kinetic rates. These reaction rates were then used to calculate the necessary chemical kinetic times. Chemical kinetic time equations for fuel, carbon monoxide and NO$_x$ were obtained for Jet-A fuel and methane with and without water injection to water mass loadings of 2/1 water to fuel. A similar correlation was also developed using data from NASA's Chemical Equilibrium Applications (CEA) code to determine the equilibrium concentrations of carbon monoxide and nitrogen oxide as functions of overall equivalence ratio, water to fuel mass ratio, pressure and temperature ($T_3$). The temperature of the gas entering the turbine ($T_4$) was also correlated as a function of the initial combustor temperature ($T_3$), equivalence ratio, water to fuel mass ratio, and pressure.

Combustion; Chemical kinetic times; Jet-A; Methane; Nitrogen oxides; Carbon monoxide

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