An ion-exchange hollow fiber is prepared by introducing into the wall of the fiber polymerizable liquid monomers, and polymerizing the monomers therein to form solid, insoluble, cross-linked, ion-exchange resin particles which embed in the wall of the fiber. Excess particles blocking the central passage or bore of the fiber are removed by forcing liquid through the fiber. The fibers have high ion-exchange capacity, a practical wall permeability and good mechanical strength even with very thin wall dimensions. Experimental investigation of bundles of ion-exchange hollow fibers attached to a header assembly have shown the fiber to be very efficient in removing counterions from solution.

8 Claims, 5 Drawing Figures
FIG. 4

FIG. 5
ION-EXCHANGE HOLLOW FIBERS

ORIGIN OF THE INVENTION

The invention described herein was made in the performance of work under a NASA contract and is subject to the provisions of Section 305 of the National Aeronautics and Space Act of 1958, Public Law 83-568 (72 Stat. 435; 42 USC 2457).


BACKGROUND OF THE INVENTION

1. Field of the Invention

The present invention relates to ion-exchange materials and, more particularly, this invention relates to novel ion-exchange hollow fibers and methods of forming and using such fibers.

2. History of the Prior Art

Current technology for removal of ions from dilute streams is largely oriented to the use of conventional packed, ion-exchange beds. These processes, however, have their problems. There is, for example, significant current effort toward the development of macroreticular: porous in the ion-exchange beads which would be less susceptible to irreversible clogging. There are problems in the preparation of beads which have adequate porosity but which are still not unduly fragile. In the preparation of commercial ion-exchange beads, the process is as follows:

A cross-linked polymer bead is formed by reacting, for example, styrene and divinylbenzene. The percentage of cross-linker (divinylbenzene) determines the extent of swelling in the final bead as ions are exchanged. The greater the percentage of cross-linker, the less the swelling. Concurrently, the greater the level of cross-linking, the slower will be the diffusion of exchanging ions into and out of the beads, and the slower will be the process.

After the bead is formed, a chemical reaction such as sulfonation or chloromethylation is used to form the ion-exchange sites. From the description it is apparent that there helps conflicting demands: high cross-link density helps stability but reduces product rate. Similarly, high ion-exchange capacity from the second step induces large swelling excursions, but provides greater capacity. Swelling of the resin beads occurs due to the osmotic pressures which are generated when the beads are exposed to different concentrations of various electrolytes. Pressure drop build-up is irregular and troublesome in regeneration processes. The choice of operating cycles is not straightforward at all and the beads are not inexpensive.

An alternative exists in semipermeable flat membranes but the technology is still in its infancy and the costs to efficiency ratio of membrane processes is not very satisfactory. Ion-exchange membranes offer significant advantages in separation processes with respect to ion-exchange resin beads. When the ion-exchange resins are in the form of membranes, they can be in contact with the solution to be separated and the stripping solution simultaneously and the ion-exchange process can be continuous rather than cyclic.

A continuous process with ion-exchange membranes has several advantages over an ion-exchange resin column. Some of these are: (1) a separate regeneration step is eliminated, (2) the product has a constant compositi

tion, (3) the hold-up volume is low, and (4) operation and control of the process are simple. The controlling factors with respect to rate of transport of ions across such membranes are as follows: (a) The membrane must be thin so that the resistance to diffusion is minimal. (b) The membrane must be semipermeable, i.e. if the membrane has positive charges they must be so uniformly and closely distributed that any solid trying to permeate the structure will always see some of the fixed charges. The consequence of this distribution is that only ions opposite in charge to the fixed charges can permeate the membrane. All ions having the same charges as the fixed charge are excluded by electrostatic repulsion. (c) The resistance to hydraulic permeability should be as low as possible consistent with the above requirements. Since this requirement and the preceding requirement are conflicting, all practical membranes are a compromise of these two objectives. (d) The flux of ions across all membranes is proportional to the area available for transfer.

These requirements result in certain limitations of the flat membrane system. The productivity per unit volume is unsatisfactory and the membranes must be supported in any separation device. Ion-exchange membranes cannot be manufactured by the same techniques utilized to form ion-exchange beads since the swelling resulting from the formation of the ion-exchange site is too great to be borne by membranes which have a low degree of cross-linking. However, if the degree of cross-linking is raised, the membrane is too brittle to be useful. Most flat ion-exchange membranes are formed by first forming ion-exchange beads and then milling the beads into a thermoplastic resin as a binder for the resin structure. In a more recent process, the thermoplastic resin is milled in the presence of a swelling agent which is then replaced with a graftable ionic monomer. After grafting, the ionic site is bound to the membrane. The mechanical requirements are satisfied by using relatively thick sheets, in the range of 100–300 microns.

The hollow fiber configuration of a membrane offers the opportunity to prepare thin-walled devices with very large surface areas. Such a device would also provide the flexibility of high transport rates per unit volume and the possibility of continuous operation without the need for regeneration cycles. Further advantage over other configurations is that supports are not required for the hollow fibers.

Anionic exchange hollow fibers have not been reported. Sulfonic acid cationic exchange type of hollow fibers have been prepared by irradiating polyethylene hollow fibers, immersing the irradiated fibers in styrene and heating the mixture to effect grafting. The fibers are then swollen in dichloromethane and sulfonated with chlorosulfonic acid, followed by hydrolysis. This procedure requires several steps, effects a random ion-exchange capacity and is limited to special reactants. Post-treatment of hollow fibers is further limited since the very small cross-section of the fibers and the fine porosity of the walls prevents introduction of preformed polymers into the bore of impregnation into the walls.

SUMMARY OF THE INVENTION

Ion-exchange hollow fibers are prepared according to the invention by introducing into the wall of the preformed fiber polymerizable liquid monomers and polymerizing the monomers therein to form solid, insol-
uble, ion-exchange resin particles embedded within the wall of the fiber.

In a preferred embodiment of the invention, the mixture of polymerizable monomers is fed into one end of the fiber and pumped through the other end. In a prepolymerizable mixture, the fiber will also exude through the pores in the walls. After the particles of polymer have been formed, the excess particles blocking the central passage are removed by forcing liquid through the bore of the fiber.

The ion-exchange fibers of the invention have several further advantages as compared to resin beads or flat membranes. There is reduced pressure drop available through the controlled dimensions of the fluid passages. In addition, the wall thickness which holds the ion-exchange sites can be reduced to 30-40 microns without prejudicing the integrity of the structure, such as is experienced with reticulated beads. Finally, the difficulties experienced by the swelling and deswelling of spheres is reduced in the cylindrical configuration of the fiber. This, in turn, allows more rapid unloading of the active sites during the regeneration cycle.

The ion-exchange hollow fibers of the invention have a high ion-exchange capacity, good mechanical strength and uniform ion-exchange capacity throughout. The treated fibers will find use as membranes in water treatment, dialysis and generally to separate ionic solutions. In the case of constraining cationic resins such as polyelectrolyte particles within the pores of the walls, the fibers can be utilized to separate chromate, phosphate, uranate, uranium sulfate complexes, cyanide or carbonate from solutions in industrial applications such as waste water treatment, process streams, plating baths or mine wastes. Hollow fibers treated to contain anionic resins such as sulfonic resins can be utilized to remove cations from aqueous solution.

A very attractive, continuous ion removal process is provided with the fibers of the invention utilizing the Donnan principle. It has been shown that when a dilute polyvalent ion is separated by an ion selective membrane from a concentrated solution of a lower valency, the mobile ions will distribute between the two solutions in such a way that the multivalent ion will concentrate in the more concentrated solution.

One very important application of such a process is in chromate inhibition. Chromates are used as corrosion inhibitors in various industrial water streams such as in boilers and cooling water service. In time, these streams build up in calcium and other cations which precipitate with carbonate. Periodically, a purge is required. This chromate contaminated blowdown is a toxicity hazard and cannot be disposed of indiscriminately. The disadvantages of packed ion-exchange resin beads have been discussed. An alternative does exist in the precipitation of a reduced chromate. This is effective for processing various liquors but is not economically attractive and not easily manageable for streams having low chromate concentration levels.

In accordance with the invention, the Donnan pumping principle is extremely effective using strong Cl⁻ concentrations to pump dilute chromate against its concentration gradient.

Thus, when a K₂CrO₄ solution of 10 ppm is fed through the core of the polyelectrolyte resin treated fibers of the invention, and the outside of the fibers bathed with a more concentrated NaCl solution, the exit concentration of CrO₄⁻ was found to be less than 0.02 ppm. The permeability of the wall, the charge density of the ion exchanger, the relative velocities of the dilute and concentrated streams were responsible for the separation and concentration effected.

These and other advantages of the invention will become apparent as the invention becomes better understood by reference to the following detailed description when considered in conjunction with the following drawings.

**BRIEF DESCRIPTION OF THE DRAWINGS**

FIG. 1 is a cross-sectional view of a single hollow fiber;

FIG. 2 is a schematic view of a hollow fiber tube assembly;

FIG. 3 is a sectional view along line 3—3 of FIG. 2;

FIG. 4 is a graph showing the amount of Cl⁻ and Cr⁺⁶ transferred across the wall of an ion-exchange hollow fiber; and

FIG. 5 is a graph showing the product concentrations of Cl⁻ and Cr⁻⁶ in the feed and effluent as a result of separation across an ion-exchange hollow fiber.

**DESCRIPTION OF THE PREFERRED EMBODIMENTS**

Referring now to FIG. 1, a hollow fiber 10 has an annular configuration, typically cylindrical. The wall 12 of the fiber 10 includes an outer skin 14 and an inner skin 16. The wall 12 contains microvoids or pores 18 which join across the wall to form a path or channel providing microporosity or semipermeability characteristics to the fiber. Particles 20 of ion-exchange resin are formed within and embedded within the pores without blocking the pores in accordance with the invention.

The dimensions of the fibers are selected depending on the use and capacity of the end device. Length of the fibers is the least important dimension and is dependent on the amount of ion to be removed, the flux rate through the wall, the flow rate of solution, the number of fibers in a bundle, and the practical length of the unit. The wall thickness, permeability and ion-exchange capacity are the most influential factors in the performance and effectiveness of ion-exchange hollow fibers.

The wall thickness is preferably as thin as practical under the circumstances and is preferably as low as 10 microns, typically from 20 to 300 microns, preferably 30-100 microns. The inside diameter is always maintained as low as possible to maximize available surface area. The inside diameter typically ranges from 50 to 1000 microns.

The morphology of the fiber used for resin deposition will determine the pore density and average pore size of the resulting ion-exchange fibers. Characterization of membrane morphology is in itself a difficult study, so that phenomonological parameters have been used by many investigators to describe membrane properties.

The descriptors found most useful are the water content of the membrane, the hydraulic resistance of the membrane, and the permeability rate of several solutes through the membrane walls. These values can sometimes be reconciled with a pore model of transport, when an independent measurement of either pore area or pore diameter can be achieved.

Pore diameters (d) of the largest pores present can be measured by the pressure of air required to displace a fluid of known surface tension from the walls of the pores if the fluid wets the pore wall. For isotropic structures, the relationship between air pressure P (in psi) and surface tension γ (in dynes/cm) is given by:
Atmosphere pressure, measured at the wall are from there atmosphere pressure level that the fibers can contain, were unsuccessful. This indicates that the average pore size is smaller than 1000 angstroms.

The hydraulic permeability of the fiber bundles can be measured by determining the rate of water permeated per unit area (A) under a specified driving force. The relationship

$$J_w = L_P A \Delta P.$$  

(1)

has been found linear over a range of pressures. $J_w$ is measured in cc/sec, and $\Delta P$ is expressed in atmospheres.

The dimensions of $L_P$ are then cm$^2$ cm$^{-2}$ sec$^{-1}$ atm$^{-1}$. Another useful expression of this relationship is in terms of the inverse of the permeability coefficient, i.e., $1/L_P$. The inverse can be considered a resistance coefficient $R_{HP}$; when membranes are used in series, the total resistance is an additive function of the individual resistances.

A third sensitive descriptor of membrane structure is formed by extruding the molten polymer or polymer solution through an annular orifice and allowing fiber formation in dimethyl acetamide, dimethyl formamide or a volatile solvent into a liquid coagulant; dry spinning (spinning from a polymer solution into a liquid coagulant); or by melt spinning. The tubular cross-section is formed by extruding the molten polymer or polymer solution through an annular dye or spinneret.

The fibers should preferably be formed from high-strength, high-modulus fibers since these can be formed with thin walls, yet allow higher operating pressure drops and permit negation of osmotic flows accompanying the ion-exchange phenomenon. The fibers should also exhibit chemical compatibility with and stability to the ion-exchange polymerization system. The polymerization system should not dissolve or permanently alter the fiber morphology during formation and embedment of the ion-exchange particles.

Textile grade acrylic fibers, suitably polymers of acrylonitrile, have been found to be compatible with the polyquaternary polymer reactants of this invention. The polyacrylonitrile may be homopolymers or copolymers containing up to 20% by weight of comonomers such as ethyl acrylate or vinyl acetate.

The acrylonitrile polymers are wet spun from a 15–25% solids solution in a highly polar solvent such as dimethyl acetamide, dimethyl formamide or 40% CaSCN through an annular spinneret into a liquid coagulant such as water. The pore size and distribution can be controlled by selection of solvent. A specific example follows.

**EXAMPLE 1**

A 20% solids solution of fiber grade polyacrylonitrile in dimethyl acetamide was wet spun through an annular spinneret and coagulated in a bath of water to form lengths of hollow fiber having an I. D. of 0.0250 cm and a wall thickness of 0.0040 cm. The diffusivities of certain solutes through the wall of the fiber of Example 1 are provided in the following table.

<table>
<thead>
<tr>
<th>Solute</th>
<th>Diffusivity (cm$^2$/sec)</th>
</tr>
</thead>
<tbody>
<tr>
<td>Creatinine</td>
<td>$6 \times 10^{-5}$</td>
</tr>
<tr>
<td>Glucose</td>
<td>$100 \min/cm$</td>
</tr>
</tbody>
</table>

The ion-exchange resin forming reactants are capable of penetrating the pores of the wall of the hollow fiber.
The reactants are also preferably capable of entering the bore of the fiber and being pumped through the fiber and exuding through the walls. The reactants therefore are selected to have a molecular weight less than 2,000 to 6,000 and are capable of forming cross-linked insoluble particles within the pores of the fiber wall. Quaternized, cross-linked, insoluble copolymers of unsubstituted and substituted vinyl pyridines and a dihalo organic compound are spontaneously formed at ambient temperature on mixing the two monomers in bulk, in solution or in suspension in solvent or in water suspension. The reaction proceeds spontaneously in bulk, in solution or in suspension as disclosed in pending application Ser. No. 102,239, the disclosure of which is incorporated herein by reference. The amount of cross-linking may be varied according to the composition and reaction conditions. The reaction yields small particles capable of depositing within the pores without blocking the pores. The non-embedded particles can readily be washed out of the fiber.

The polyquaternary, water insoluble, cross-linked materials are prepared by reacting a vinyl pyridine with a dihalo organic compound of the formula:

\[
X-R_1-X
\]

where X is halo, preferably bromo, chloro or iodo and R₁ is a divalent organic radical such as alkylene, alkylene, alkynylene, arylene, alkarylene or aralkylene. R₁ may also be alkylthioalkylene or alkylalkoxyalkylene of the formula \((\text{CH}_2)_n(\text{Z})_m(\text{CH}_2)_o\), where Z is oxygen or sulfur and x, y and z are integers from 1 to 100. R₁ may also be of prepolymer or polymeric length of up to 6,000 molecular weight such as a bromo-terminated polybutadiene, but, preferably has a carbon content of from 1 to 20 carbon atoms to provide an increased charge center density per unit volume and weight of the polymeric product. R₁ may be substituted with other groups that do not interfere with the polymerization reaction or properties of the polymer such as hydroxyl, alkyl, aryl, nitro, cyano or similar groups.

Representative dihalo organic compounds are 2,4-alkylene or alkynylene halides such as dibromo methane, 1,2-dibromoethane, 1,3-dibromopropane, 1,4-dibromobutane, 1,4-dibromo-2-butene, 1,4-dibromo-2,3-dihydroxy butane, 1,5-dibromopentane, 1,6-dibromohexane, 1,8-dibromooctane, 1,10-dibromodecane, and 1,16-dibromohexadecane. The alkylene compounds such as 1,4-dibromobutene are found to be more reactive than the corresponding saturated compounds. Dihaloaromatics such as o, m and p-dichloro- or o, m and p-dibromobenzene are also suitable. Cross-linked, insoluble products would also be formed from terminally brominated prepolymers such as polyethylene, polypropylene, polybutylene, polybutadiene, polyoxoethylene, etc. As the number of carbon atoms in the dihalide increases, elastomeric properties are favored and polyelectrolyte properties decrease.

4-vinyl pyridine is the most reactive of the vinyl pyridine isomers. However, 2-methyl-5-vinyl pyridine is available at lower cost and provides products of similar properties. 2-vinyl pyridine has been found to be much less reactive than the other monomers.

The polycationic, cross-linked products are prepared simply by mixing the vinyl pyridine monomer with the dihalide in various proportions and allowing the mixture to react until solid materials are formed. The reaction is believed to proceed through a first stage in which two molecules of vinyl pyridine react with a molecule of a dibromide to form a quaternary intermediate as illustrated below:

Two molecules of the intermediate dicationic, diunsaturated cross-linking agent then react through the vinyl group to form an intermediate having a structure of the formula:

\[
\begin{align*}
\text{CH}_2 & \text{-CH}_2 \\
\text{N}^+ & \text{-R}_1 \text{-N}^+ \\
\text{CH}_2 & \text{-CH}_2
\end{align*}
\]

The intermediate reacts further to give a cross-linked network with a small amount of residual unsaturation. Although this mechanism is dominant, other intermediates also formed. The reaction proceeds spontaneously at room temperature, about 25° C., but may be accelerated by heating the reaction to a higher temperature, usually below 100° C., and suitably from 25°–60° C. The unsaturation on the growing polymer as well as on the finished resin may be utilized in further reactions, e.g. grafted onto substrates by means of Coγ radiation. Cross-linking of the product is also facilitated by irradiating the mixture with radiation capable of forming reactive species to cross-link the vinyl groups, suitably gamma radiation from a cobalt source. The reaction may be conducted in bulk, in a solvent for the monomer or in water suspension. The reaction proceeds faster in bulk, but yields are higher in solvents. Higher yields are favored in polar solvents such as dimethylsulfoxide, dimethylformamide, methanol, ethanol, or combinations thereof. Particularly high yields have been obtained with a 1/1 volume mixture of dimethylformamide and methanol. Slower reaction occurs in solvents such as benzene.

The rate of reaction is found to be much higher with bromides, as compared to the corresponding chlorides. The ratio of monomers is controlled such that there is an excess of dibromide in the mixture. A suitable ratio is a stoichiometric ratio of 2 mols of vinyl pyridine to at least 1 mol of the dibromide. It has been found that when the polymerization is conducted with an excess of vinyl pyridine, unchanged vinyl pyridine can be recovered. It has further been found that oxygen and carbon dioxide interfere, inhibit and slow the reaction. Also...
outlet header. The polymerization proceeded spontaneously for about three days. Before polymerization was complete, water was injected into the header with a syringe to open the central channel of the fibers and to prevent plugging. The polymerization proceeded within the pores of the fiber wall to form particles of an insoluble, cross-linked polyquaternary resin.

EXAMPLE 3

Example 1 was repeated substituting dibromohexane for the dibromomethane. An ion-exchange hollow fiber resulted having an ion-exchange capacity about one half that of the product of Example 2.

After formation of the polyquaternary ion-exchange resin in the fiber walls, significant changes in the transport properties of ionic species were expected. The end point of such transport was governed by the equilibrium expressions derived by Donnan, and the most significant property which governed this equilibrium was the charge density of the immobilized species. If the charge density is adequately high, the membrane will be truly "semi-permeable" and the Donnan assumptions apply. In these experiments, the immobilized species was the quaternary nitrogen, and the gegenions were negative ions. The concentrations of the fixed negative ions determined by exchanging the gegenions to (OH–) forms by thorough washing in 0.1 N NaOH. The fibers were then placed in a known excess of HCl, and after 24 hours the fibers were removed, rinsed, and the remaining HCl titrated.

Titrination of an aliquot of fiber indicated that 0.92 milliequivalents of (OH–) were bound to 1.3 gms of fiber to give 0.71 meq/gm. The sample represented 155 cm² so that the charge per unit area of membrane was 0.006 meq/cm².

The introduction of a significant weight of ion-exchange polymer into the polyacrylonitrile fiber should change the permeability of the fiber wall to water and to nonionic solutes, since the void volume is reduced. This reduction is shown by the data in Table 2.

<table>
<thead>
<tr>
<th>Table 2</th>
</tr>
</thead>
<tbody>
<tr>
<td>Hydraulic Permeability</td>
</tr>
<tr>
<td>Solute Resistances</td>
</tr>
<tr>
<td>Creatinine</td>
</tr>
<tr>
<td>Glucose</td>
</tr>
</tbody>
</table>

The hydraulic permeability has been reduced by a factor of four, and the solute resistance of creatinine has increased by a factor of five. The solute measurements were performed both initially, and after ion-exchange deposition, with a 0.86% NaCl supporting electrolyte concentration. It was later found that the ion-exchange resin operated effectively at this NaCl concentration so the permeability decrease may reflect, in part, a wall thickness increase due to osmotic swelling.

The property that is most significant, of course, is the semipermeability of the fiber wall with respect to anions. To examine this function, CrO₄²⁻ anions were used because of the ease with which very small concentrations can be measured and because of the current interest in removing such ions from water cooling systems.

The hollow fiber assembly of Example 2 was utilized to remove chromate ions from a dilute solution (10 ppm Cr⁺₆). Referring again to FIGS. 2 and 3, the bundle 24 of fibers was immersed in a vessel 36 containing a NaCl aqueous solution. The dilute chromate solution was injected under controlled pressure into the funnel 28 and a solution having very low chromat and chloride ion content
was removed from outlet header 22. A specific example follows.

 EXAMPLE 4

A concentration of $2.0 \times 10^{-4}$ M $K_2CrO_4$ was pumped at 3.0 cc/min through a fiber bundle containing 150 fibers (153.1 cm$^2$). The fiber was bathed externally in a slowly stirred solution containing initially 0.25 M NaCl. The core solution leaving the fiber bundle contained less than $4 \times 10^{-6}$ M CrO$_4^-$ (present low limit of analysis), and $6.5 \times 10^{-3}$ M Cl$^-$. The external bath increased continuously in CrO$_4^-$ concentration, and decreased in Cl$^-$ concentration as shown in FIG. 4.

In a perfectly semipermeable system, the number of milliequivalents of CrO$_4^-$ permeating the wall would be exactly equal to the number of milliequivalents of Cl$^-$ moving in the other direction. The actual situation is not perfect. FIG. 5 shows the number of milliequivalents of each ion that permeated the semipermeable membrane over a period of time. The actual situation when the feed rate was increased to 16 cc/min, the exit concentration of the fiber, the anions present exchange across the semipermeable walls of the fibers. The exit concentration of the Cl$^-$ was depleted to 0.06 M. The total volume in the external system had increased to 1,000 cc as a result of osmotic pumping of water.

Considering the mass transfer of CrO$_4^-$ in the same way that nonionic species were characterized, one can take the initial rate of transfer and calculate a membrane permeability:

$$P = \frac{\frac{3.0}{10} \times 153}{2 \times 10^{-4}} = 1.3 \times 10^{-3} \text{ cm/sec}$$

This is seen to be significantly higher than the rates attained for non-electrolytes. The comparison is, of course, not valid since the derivation assumes that only the concentration difference in (CrO$_4^-$) available and the small gradient of (CrO$_4^-$), the lack of ideality is not surprising. At the end of the experiment, when 15 liters had been processed through the fibers, the external bath was 40$\times$10$^{-4}$ M in CrO$_4^-$, and the (Cl$^-$) was depleted to 0.06 M. The total volume in the external system had increased to 1,000 cc as a result of osmotic pumping of water.

As the core solution flows through the length of each fiber, the anions present exchange across the semipermeable walls of the fibers. The exit concentration of the core solution is <0.02 ppm Cr$^{+6}$ and 22 ppm Cl$^-$. Present federal standards for effluents are less than 0.05 ppm of chromium. The external solution increases in Cr$^{+6}$ concentration, and after 13 liters of core fluid have been treated, the external solution has been enriched by 4.9 moles of Cr. The process will continue so long as the Donnan potential which drives the system toward the equilibrium:

$$N=(\text{Cl}^-)/((\text{Cl}^-)^2)=(\text{CrO}_4^-)/((\text{CrO}_4^-)^2)$$

As the core solution flows through the length of each fiber, the anions present exchange across the semipermeable walls of the fibers. The exit concentration of the core solution is <0.02 ppm Cr$^{+6}$ and 22 ppm Cl$^-$. Present federal standards for effluents are less than 0.05 ppm of chromium. The external solution increases in Cr$^{+6}$ concentration, and after 13 liters of core fluid have been treated, the external solution has been enriched by 4.9 moles of Cr. The process will continue so long as the Donnan equilibrium condition has not been achieved. When the feed rate was increased to 16 cc/min, the exit concentration of dichromate was found to be as low as 0.7 ppm.

The applications for such a process are numerous. The process is equivalent to a continuous form of the Donnan system toward the equilibrium: driving force is the Donnan potential which drives the nonionic species to a function of the concentration difference in (CrO$_4^-$), and the lack of permeating the wall would be significantly higher than the rates attained for non-electrolytes. The comparison is, of course, not valid since the derivation assumes that only the concentration difference in (CrO$_4^-$) available and the small gradient of (CrO$_4^-$), the lack of ideality is not surprising. At the end of the experiment, when 15 liters had been processed through the fibers, the external bath was 40$\times$10$^{-4}$ M in CrO$_4^-$, and the (Cl$^-$) was depleted to 0.06 M. The total volume in the external system had increased to 1,000 cc as a result of osmotic pumping of water.

At the same time it is known that from the weight of water imbibed per gram of untreated fiber, that the amount of void space in the wall is fairly large. The pores provide a relatively large volume in which the ion-exchange particles can be formed and permit the achievement of fairly high ion-exchange capacities. It is to be understood that only preferred embodiments of the invention have been described and that numerous substitutions, alterations and modifications may be made without departing from the spirit and scope of the invention as defined in the following claims.

What is claimed is:

1. A method of forming an ion-exchange medium comprising the steps of:
   - introducing into one end of the central channel of a hollow, thin-wall, microporous fiber of a synthetic organic resin having pores which join to form a path across the wall of the fiber a mixture of liquid reactive monomers until the mixture appears at the second end thereof and exudes through the wall of the fiber to the outside surface thereof; and
   - reacting the monomers within the pores to form solid, insoluble ion-exchange resin particles having a diameter smaller than the pore diameter, and having an even distribution of fixed charge, said particles being embedded and constrained within said pores without blocking said pores or said path.

2. A method according to claim 1 further including the step of removing any non-embedded particles from said pores and from the central channel of the fiber.

3. A method according to claim 1 in which the fiber has a wall thickness of from 1 to 300 microns and an internal diameter of from 50 to 1,000 microns.
4. A method according to claim 3 in which the pores are evenly distributed throughout the wall of the fiber and have an average diameter less than 1,000 angstroms.

5. A method according to claim 1 in which the monomers comprise vinyl pyridine and an organic dihalide.

6. A method according to claim 5 in which the vinyl pyridine is 4-vinyl pyridine and the organic dihalide is a compound of the formula:

\[ X - R_1 - X \]

\( X \) is halogen and \( R_1 \) is a divalent organic radical containing from 1-20 carbon atoms.

7. A method according to claim 6 in which \( X \) is chloro or bromo and \( R_1 \) is a divalent organic radical selected from alkylene, alkenylene, alkynylene, arylene, alkaryl and aralkylene.

8. A method according to claim 7 in which the organic dihalide is 1,2-dibromoethane.