



Fluorine-Based DRIE of Fused Silica

A suitable choice of process parameters enables etching of vertical side walls.

NASA's Jet Propulsion Laboratory, Pasadena, California

A process of deep reactive-ion etching (DRIE) using a fluorine-based gas mixture enhanced by induction-coupled plasma (ICP) has been demonstrated to be effective in forming high-aspect-ratio three-dimensional patterns in fused silica. The patterns are defined in part by an etch mask in the form of a thick, high-quality aluminum film. The process was developed to satisfy a need to fabricate high-aspect-ratio fused-silica resonators for vibratory microscopes, and could be used to satisfy similar requirements for fabricating other fused-silica components.

The development of the process involved manipulation of some process parameters, including the selection of gases, the mixture ratios of the gases, the process pressure, and power of the radio-frequency signal used to excite the

ICP. It was found that polymeric materials that contained silicon and/or fluorine were formed on the side walls during etching and were subsequently etched away, resulting in anisotropic etching. It was also found that helium added to gas mixtures contributes to cooling of substrates and thereby helps in forming vertical side-wall patterns.

One version of the process was found to effect etching of vertical side walls in 100- μm -deep fused silica at a rate of about 0.4 $\mu\text{m}/\text{min}$. Some of the parameters of this version of the process were the following:

- ICP power 1.2 kW;
- Reactive-ion etching power 0.5 kW;
- Process pressure 6 mtorr (≈ 0.8 Pa);
- Gas mixture of 1 part C_4F_8 to 1.6 parts of H_2 ; and
- Thickness of aluminum mask 5 μm .

The rate of etching of the fused silica was found to be about 5 times the rate of etching of the aluminum.

This work was done by Karl Yee, Kirill Shcheglov, Jian Li, and Daniel Choi of Caltech for NASA's Jet Propulsion Laboratory.

In accordance with Public Law 96-517, the contractor has elected to retain title to this invention. Inquiries concerning rights for its commercial use should be addressed to:

*Innovative Technology Assets Management
JPL*

Mail Stop 202-233

4800 Oak Grove Drive

Pasadena, CA 91109-8099

(818) 354-2240

E-mail: iaoffice@jpl.nasa.gov

Refer to NPO-43837, volume and number of this NASA Tech Briefs issue, and the page number.

Mechanical Alloying for Making Thermoelectric Compounds

Constituents are ball-milled into a powder, which is then hot pressed.

NASA's Jet Propulsion Laboratory, Pasadena, California

An economical room-temperature mechanical alloying process has been shown to be an effective means of making a homogeneous powder that can be hot-pressed to synthesize a thermoelectric material having reproducible chemical composition. The thermoelectric materials to which the technique has thus far been applied with success include rare-earth chalcogenides [$\text{La}_{3-x}\text{Te}_4$ ($0 < x < 0.33$) and $\text{La}_{3-x}\text{Yb}_y\text{Te}_4$ ($0 < x < 1, 0 < y < 1$)] and Zintl compounds (including $\text{Yb}_{14}\text{MnSb}_{11}$ and $\text{Yb}_{14}\text{BiSb}_{11}$). The synthesis of a given material consists of the room-temperature thermomechanical-alloying process followed by a hot-pressing process. Relative to synthesis of nominally the same material by a traditional process that includes hot melting, this synthesis is simpler and yields a material having superior thermoelectric properties.

The room-temperature mechanical alloying process is, more specifically, a

ball-milling process. It begins inside an argon-filled glove box, wherein elemental constituents in amounts corresponding to their desired proportions in the thermoelectric material to be synthesized are loaded into a vial that contains milling balls. The vial and milling balls are made of a material compatible with the material to be synthesized. (For synthesizing $\text{Yb}_{14}\text{MnSb}_{11}$, one uses a vial and balls made of tungsten carbide; for synthesizing $\text{La}_{3-x}\text{Te}_4$ or $\text{La}_{3-x}\text{Yb}_y\text{Te}_4$, one uses a vial and balls made of stainless steel.) Next, the filled vial is removed from the glove box and clamped onto a commercially available mixer/mill machine, which is used to shake the vial for as long as 40 hours to effect ball milling.

After ball milling, the vial is returned to the glove box, wherein the powder produced by the ball milling is loaded into a graphite die for hot pressing.

- In the case of $\text{Yb}_{14}\text{MnSb}_{11}$, it is neces-

sary to sandwich the powder between two graphite foil layers at each end. In ascending order, the resulting assembly inside the die consists of one or more spacer(s), two graphite foil layers, the powder, two more graphite foil layers, and a plunger that presses down on the aforementioned components.

- In the case of $\text{La}_{3-x}\text{Te}_4$ or $\text{La}_{3-x}\text{Yb}_y\text{Te}_4$, the plunger is made of graphite, the inside of the die is lined with graphite foil, and the powder touches the top and bottom spacers, which are coated with boron nitride to prevent adhesion.

The die and its contents are then placed in a hot press, wherein the powder is subjected to a temperature-vs.-time and a pressure- (or force)-vs.-time profile, specified for the material to be synthesized (for example, see figure), to consolidate the powder into a solid mass of requisite density. After this hot pressing, the mass is removed from the die. In the case of $\text{La}_{3-x}\text{Te}_4$ or $\text{La}_{3-x}\text{Yb}_y\text{Te}_4$, the