The present invention is directed to the effective dispersion of carbon nanotubes (CNTs) into polymer matrices. The nanocomposites are prepared using polymer matrices and exhibit a unique combination of properties, most notably, high retention of optical transparency in the visible range (i.e., 400-800 μm), electrical conductivity, and high thermal stability. By appropriate selection of the matrix resin, additional properties such as vacuum ultraviolet radiation resistance, atomic oxygen resistance, high glass transition (Tg) temperatures, and excellent toughness can be attained. The resulting nanocomposites can be used to fabricate or formulate a variety of articles such as coatings on a variety of substrates, films, foams, fibers, threads, adhesives and fiber coated prepreg.

The properties of the nanocomposites can be adjusted by selection of the polymer matrix and CNT to fabricate articles that possess high optical transparency and antistatic behavior.

6 Claims, 10 Drawing Sheets
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\[ \text{H}_2\text{N-} \text{Ar} - \text{H}_2\text{N} + \text{O} \text{Ar}^* \text{O} + \text{CNT} \]

\[ \downarrow \text{23°C, N}_2 \text{Polar Aprotic Solvent} \]

\[ \text{HOOC} \text{ COOH} \]

\[ \text{-H}_2\text{O} \]

\[ \text{FIG. 1} \]

\[ \text{HO - Ar}^*\text{-- OH + X-- Ar}^{*\prime}\text{--X + CNT} \]

\[ \text{-H}_2\text{O} \text{ inert atmosphere Alkali metal base Polar aprotic solvent Heat} \]

\[ \text{FIG. 2} \]
+ Laser ablated carbon nanotubes

\[ \text{+ Laser ablated carbon nanotubes} \]

\[ \text{+ Laser ablated carbon nanotubes} \]

\[ \text{+ Laser ablated carbon nanotubes} \]

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Chemical vapor deposition carbon nanotubes

\[
\text{O} = \text{C} = \text{O}
\]

\[
\text{O} = \text{C} = \text{O} \quad \text{N} = \text{O} \quad \text{HO} \quad \text{O}
\]

\[
\text{NH}_2 + \text{Ph}_2\text{Ph} \quad 23^\circ\text{C}, \text{N}_2 \quad \text{Polar Aprotic Solvent} \quad \text{H}_2\text{O} \quad \text{H}^+ \quad \text{FIG. 7}
\]
+ Chemical vapor deposition carbon nanotubes

+ Chemical vapor deposition carbon nanotubes

+ Chemical vapor deposition carbon nanotubes

+ Chemical vapor deposition carbon nanotubes

\[ \text{FIG. 8} \]
+ Chemical vapor deposition carbon nanotubes

\[
\begin{align*}
\text{H}_2\text{N} & \quad \text{NH}_2 \\
\text{O} & \quad \text{O} \\
\text{O} & \quad \text{O} \\
\text{O} & \quad \text{O} \\
\text{Ph} & \quad \text{Ph} \\
\text{Ph} & \quad \text{Ph} \\
\text{Ph} & \quad \text{Ph} \\
\text{Ph} & \quad \text{Ph} \\
\text{23°C, N}_2 & \quad \text{Polar Aprotic Solvent} \\
\text{H}_2\text{O} & \quad \text{-H}_2\text{O} \\
\end{align*}
\]
+ Chemical vapor deposition carbon nanotubes

\[
\text{ PIC 10 }
\]

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\text{ PIC 10 }
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\text{ PIC 10 }
\]
ELECTRICALLY CONDUCTIVE, OPTICALLY CLEAR POLYMER/CARBON NANOTUBE COMPOSITES

This application is a divisional application of commonly-owned patent application Ser. No. 10/288,797, filed Nov. 1, 2002, issued as U.S. Pat. No. 7,588,699 which, pursuant to 35 U.S.C. §119, claimed the benefit of priority from provisional patent application having U.S. Ser. No. 60/336,109, filed on Nov. 2, 2001, the contents of which are incorporated herein in their entirety.

ORIGIN OF INVENTION

The invention described herein was jointly made by employees of the U.S. Government, contract employees and employees of the National Research Council, and may be manufactured and used by or for the government for governmental purposes without the payment of royalties thereon or therefor.

BACKGROUND OF THE INVENTION

1. Field of the Invention

The present invention is directed to methods of preparation that effectively disperse carbon nanotubes into polymer matrices, and the novel nanocomposites that result therefrom.

2. Description of the Related Art

Since carbon nanotubes (CNTs) were discovered in 1991 (S. Iijima, Nature 354, 56, 1991), significant interest has been generated due to their intrinsic mechanical, electrical, and thermal properties (P. M. Ajayan, Chem. Rev. 99, 1787, 1999). Early studies focused on CNT synthesis and theoretical predictions of physical properties. Due to the recent development of efficient CNT synthesis (A. Thees et al., Science 273, 483, 1996) and purification procedures (A. G. Rinzler et al., Appl. Phys. A 67, 29, 1998), some applications have been realized. However, these applications have relied on the use of pure CNTs, not nanocomposites. Examples include a carbon nanotube in scanning probe microscopy (S. S. Wong et al., J. Am. Chem. Soc. 120, 603, 1998), single wall carbon nanotube (SWNT) transistor (S. J. Tans et al., Nature 393, 49, 1998), and field emission display (Q. H. Wang et al., Appl. Phys. Lett. 70, 5308, 1997). There have been few reports on the development of nanocomposites using CNTs as reinforcing inclusions in a polymer matrix primarily because of the difficulty in dispersing the nanotubes. This difficulty is partially due to the non-reactive surface of the CNT. A number of studies have concentrated on the dispersion of CNTs, but complete dispersion of the CNTs in a polymer matrix has been elusive due to the intrinsically strong van der Waals attraction between adjacent tubes. In practice, attempts to disperse CNTs into a polymer matrix lead to incorporation of agglomerates and/or bundles of nanotubes that are micron sized in thickness and, consequently, they do not provide the desired and/or predicted properties. Most of the dispersion related studies have focused on modifying the CNT surface chemistry. Many researchers have studied the functionalization of CNT walls and ends. One example is fluorination of CNT surfaces (E. T. Mickelson et al., J. Phys. Chem. 103, 4318, 1999), which can subsequently be replaced by an alkyl group to improve the solubility in an organic solvent. Although many researchers have tried to functionalize CNT ends and exterior walls (as a means to increase solubility) by various approaches such as electrochemistry and wrapping with a functionalized polymer, the solubility of these modified tubes was very limited. Other methods of CNT modification include acid treatment (i.e. oxidation) and use of surfactants as a means of improving solubility and compatibility with organic polymers. It has been noted that modifications of the nanotube chemical structure may lead to changes in intrinsic properties such as electrical conductivity (X. Gong et al., Chem. Mater. 12, 1049, 2000). Ultrasonic treatment has also been used as a means to disperse CNTs in a solvent. Upon removal of the sonic force, the tubes agglomerate and settle to the bottom of the liquid.

Individual SWNTs can exhibit electrical conductivity ranging from semi-conductor to metallic depending on their chirality, while the density is in the same range of most organic polymers (1.33x10^14 g/cm^3). In the bulk, they form a pseudo-metal with a conductivity of approximately 10^5 S/cm (Kaiser et al., Physics Reviews B, 57, 1418 1998). The conductive CNTs have been used as conductive fillers in a polymer matrix to enhance conductivity, however the resulting nanocomposites exhibited little or no transparency in the visible range (400-800 nm)(P. M. Ajayan, (Physical Review B, 58, R7492, 1998) and Curran et al., (Advanced Materials, 10, 1091, 1998) reported conjugated polymer-CNT composites using multi-wall CNTs, which showed that the percolation concentration of the CNTs exceeded 5 wt %. The resulting nanocomposites were black with no transparency in the visible region. Shaffer and Windle (Advanced Materials, 11, 937, 1999) reported conductivity of a multi-wall CNT/poly (vinyl alcohol) composite, which also showed percolation above 5 wt % nanotube loading and produced a black nanocomposite. The same group (J. Sandler, M. S. P. Shaffer, T. Prasse, W. Bauhofer, K. Schulte, and A. H. Windle, Polymer 40, 5967, 1999) reported another multi-wall CNT composite with an epoxy, which achieved percolation below 0.04 wt %. An optical micrograph of the CNT/epoxy composite was reported, which revealed that the CNT phase was separated from the epoxy resin, showing several millimeters of resin-rich domains. The dispersion of CNTs in this material was very poor. This agglomeration of CNTs in selected areas in the composite could explain the high conductivity observed since it provides the “shortest path” for the current to travel. Preliminary measurements of the conductivity of a CNT/poly (methyl methacrylate) (PMMA) composite were measured on a fiber (R. Haggenmueller, H. H. Gommaens, A. G. Rinzler, J. E. Fischer, and K. I. Winney, Chemical Physics Letters, 330, 219, 2000). The level of conductivity was relatively high (1.18x10^-1 S/cm) at 1.3 wt % SWNT loading. However, the optical transparency in the visible range was not determined for the fiber sample. The mechanical properties of these fibers were much less than the predicted value, which implies that the CNTs were not fully dispersed.

The present invention is directed to methods of preparation that overcome the shortcomings previously experienced with the dispersion of CNTs in polymer matrices and the novel compositions of matter produced therefrom. The resulting nanocomposites exhibit electrical conductivity, improved mechanical properties, and thermal stability with high retention of optical transparency in the visible range.

SUMMARY OF THE INVENTION

Based on what has been stated above, it is an objective of the present invention to effectively disperse CNTs into polymer matrices. It is a further objective to prepare novel polymer-CNT nanocomposites and articles derived therefrom.
Methods of preparation that were evaluated include: (1) low shear mixing of a polymer solution with CNTs dispersed in an organic solvent; (2) high shear mixing (e.g., homogenizer or fluidizer) of a polymer solution with CNTs dispersed in an organic solvent; (3) ultrasonic mixing (e.g., sonic horn at 20-30 kHz for 1-10 minutes) of a polymer solution with CNTs dispersed in an organic solvent; (4) high shear mixing (e.g., homogenizer, fluidizer, or high speed mechanical stirrer) of a polymer solution with CNTs dispersed in an organic solvent with subsequent ultrasonic mixing (e.g., sonic horn at 20-30 kHz for 1-10 minutes); (5) synthesis of the polymer in the presence of pre-dispersed CNTs; and (6) synthesis of the polymer in the presence of pre-dispersed CNTs with simultaneous sonication (e.g., 40-60 kHz in a water bath) throughout the entire synthesis process. Methods (4), (5), and (6) are applicable to a variety of polymers that can be synthesized in a solvent in the presence of the CNTs.

The resulting polymer/CNT materials exhibit a unique combination of properties that make them useful in a variety of aerospace and terrestrial applications, primarily because of their combination of improved mechanical properties, thermal stability, electrical conductivity, and high optical transmission. Examples of space applications include thin film membranes on antennas, second-surface mirrors, thermal protective coatings on satellites, and optical coatings, and multi-layer thermal insulation (MLI) blankets that are used. For these applications, materials that do not build up electrical charge are preferred. In addition to exhibiting electrical conductivity, some of these space applications also require that the materials have low solar absorptivity and high thermal emissivity. Terrestrial applications include electrically conductive coatings on a variety of substrates, electrostatic dissipative coatings on electromagnetic displays, coatings for use in luminescent diodes, antistatic fabrics, foams, fibers, threads, clothing, carpeting and other broad goods.

BRIEF DESCRIPTION OF THE FIGURES

FIG. 1 illustrates preparation of an aromatic poly(amide ether)/CNT nanocomposite.

FIG. 2 illustrates preparation of an aromatic poly(arylene ether)/CNT nanocomposite.

FIG. 3 illustrates preparation of a 0.1 wt % CNT/polyimide nanocomposite from 1,3-bis(3-aminoophenoxo) benzene (APB) and 4,4'-perfluorophenylsulfone diimide dianhydride (6FDA).

FIG. 4 illustrates preparation of a 0.1 wt % SWNT/polyimide nanocomposite from 2,6-bis(3-aminoophenoxo) benzonitrile ([6-CN]APB) and 3,3’,4,4’-oxydipthalic diimide dianhydride (ODPA).

FIG. 5 illustrates preparation of a 0.1% wt/wt LA-NT/polyimide nanocomposite from 2.4-bis(3-aminoophenoxo) phenyldiphenylphosphate oxide (APB-PPO) and ODPA.

FIG. 6 illustrates preparation of a 0.2% wt/wt LA-NT/polyimide nanocomposite from APB-PPO and ODPA.

FIG. 7 illustrates preparation of a 0.1% wt/wt CVD-NT/1-polyimide nanocomposite from APB-PPO and ODPA.

FIG. 8 illustrates preparation of a 0.2% wt/wt CVD-NT/1-polyimide nanocomposite from APB-PPO and ODPA.

FIG. 9 illustrates preparation of a 0.1% wt/wt CVD-NT/2-polyimide nanocomposite from APB-PPO and ODPA.

FIG. 10 illustrates preparation of a 0.2% wt/wt CVD-NT/2/polyimide nanocomposite from APB-PPO and ODPA.

FIG. 11 illustrates preparation of a 0.1% wt/wt LA-NT/poly(arylene ether)/SWNT nanocomposite.

DETAILED DESCRIPTION OF THE INVENTION

The present invention involves the preparation of polymer/CNT composites with a unique combination of properties.
fluidizer, or high-speed mechanical stirrer) and was subsequently placed in a glass vessel and immersed in an ultrasonic water bath operating at 40-60 kHz for several 1-10 hours to achieve initial dispersion.

Selection of Polymers

Predominantly aromatic and conjugated polymers are generally preferred for use in the preparation of polymer/CNT nanocomposites for long-term aerospace applications owing to their high-temperature resistance and high durabilities. Representative aromatic polymers and copolymers, representing the poly(amide acid), polyimide and poly(arylene ether) families, were selected based upon their solubility in several polar aprotic solvents of choice and their ability to be synthesized in the presence of the CNTs without any deleterious effects on molecular weight build-up as evidenced by a noticeable increase in solution viscosity. In some cases, target polymers with polar groups such as carboxyl, cyano, phosphine oxide, sulfone and others or conjugated polymers were selected to provide additional compatibility with CNTs. In some cases, polymers with very high optical transmission (i.e., greater than approximately 85%) at 500 nm were selected to demonstrate this approach. Particularly good results, with respect to degree of dispersion, were obtained with aromatic polymers containing polar groups.

Methods of Preparation of Composites

Several methods of preparing polymer/CNT composites were evaluated and are described in detail below.

Method (1) (Low Shear Mixing)

Low shear mixing of a pre-synthesized high molecular weight aromatic polymer solution with CNTs dispersed in an organic solvent was conducted by preparing a polymer solution in a solvent and subsequently adding the CNT dispersion (prepared as previously described). A mechanical stirrer was used to mix the two components. This approach typically resulted in poor mixing and poor dispersion. The CNTs separated from solution upon removal of the mechanical agitation. The resulting film and/or coating were black in color and exhibited poor retention of optical transmission (i.e., less than approximately 35% retention of optical transmission) at 500 nm. Optical microscopic examination of the nanocomposite film showed the presence of agglomerates of CNT bundles indicating poor dispersion.

Method (2) (High Shear Mixing)

High shear mixing (e.g., using homogenizer, fluidizer, or high-speed mechanical stirrer) of a pre-synthesized high molecular weight aromatic polymer solution with CNTs dispersed in an organic solvent was conducted by preparing a polymer solution in a solvent and subsequently adding the CNT dispersion (prepared as previously described). A flat bottom generator equipped with a homogenizer operating at about 7500 revolutions per minute (rpm) was used for approximately 20 minutes to mix the two components. Experiments were undertaken to study the effect of homogenization time on level of dispersion. Longer homogenization times (>1 hour) did not provide significant improvement in mixing and dispersion as compared to shorter times (<1 hour). This approach typically resulted in better mixing and dispersion as compared to Method (1), but the resulting nanocomposite films and/or coatings were black and exhibited poor retention of optical transmission (i.e., less than approximately 35% retention of optical transmission) at 500 nm. Optical microscopic examination of the nanocomposite film showed the presence of agglomerates of CNT bundles indicating poor dispersion.

Method (3) (Ultrasonic Mixing with Sonic Horn)

Ultrasonic mixing of a pre-synthesized high molecular weight aromatic polymer solution with CNTs dispersed in an organic solvent was conducted by preparing a polymer solution in a solvent and subsequently adding the CNT dispersion (prepared as previously described). A high power sonic horn equipped with a 13 mm probe operating at 20 kHz was used to mix the two components. Experiments were undertaken to study the effect of ultrasonic treatment time on level of dispersion. Longer ultrasonic treatment times (>10 min.) did not provide significant improvement in mixing and dispersion as compared to shorter ultrasonic treatment times (<10 min.). This high power ultrasonic treatment appeared to cause significant damage to the polymer as evidenced by a noticeable decrease in solution viscosity. This observation suggests that chemical bond cleavage is occurring that subsequently leads to a reduction in molecular weight. The possibility also exists that this high power ultrasonic treatment may cause damage (i.e., introduction of defect sites through carbon-carbon bond cleavage) to the CNTs. Modification of the chemical structure of CNTs is known to cause bulk property changes, thus this method was deemed undesirable. Nanocomposite films and/or coatings prepared from solutions that received relatively short exposures (<10 min.) to the high power sonic horn treatment exhibited improvements in electrical conductivity of 10-12 orders of magnitude; however the nanocomposite films and/or coatings exhibited moderate retention of optical transparency (i.e., 35-50% retention of optical transmission) in the visible range. Optical microscopic examination of the nanocomposite film showed the presence of agglomerates of CNT bundles indicating poor dispersion. Based on a qualitative assessment, the nanocomposite film prepared via this method exhibited marginally improved dispersion relative to the nanocomposite films prepared via Methods (1) and (2).

Method (4) (High Shear and Ultrasonic Mixing Using Sonic Horn)

A combination of high shear mixing and ultrasonic treatment was conducted by initially preparing an aromatic polymer solution in a solvent and subsequently adding the CNT dispersion (prepared as previously described). A homogenizer was subsequently used to mix the dispersion, followed by ultrasonic treatment with a high power sonic horn operated at 20 kHz. The times of each treatment were varied, but no significant differences in dispersion were apparent. This combination treatment generally gave better dispersion than one single component mixing. Nanocomposite films and/or coatings with 0.1 wt% CNT exhibited improvements in electrical conductivity of 10-12 orders of magnitude compared to a pristine polymer film. However, the nanocomposite films and/or coatings exhibited moderate retention of optical transparency (i.e., 35-50% retention of optical transmission) at 500 nm. Optical microscopic examination of the nanocomposite film showed the presence of agglomerates of CNT bundles, indicating poor dispersion. Based on a qualitative assessment, the nanocomposite film prepared via this method exhibited marginally improved dispersion relative to the nanocomposite films prepared via Methods (1) and (2).

Method (5) (Synthesis of the Polymer in the Presence of Pre-Dispersed CNTs)

Synthesis of an aromatic polymer in the presence of the CNTs was conducted by pre-dispersing the CNTs in the solvent of interest and subsequently adding the monomers. In the case of the poly(amide acid)s and polyimides, the diamine component was added first to the predispersed CNTs and allowed to be stirred until dissolved. The dihydride component was subsequently added as a solid and the progression of the polymerization was readily observable by a significant build-up in solution viscosity. The re-aggregation among the CNTs is inhibited and/or minimized due to the high viscosity of the solution, which preserves the state of CNT disper-
sion during further required processing. The polymerization was allowed to proceed under conditions analogous to those generally used for the particular polymer type using a mechanical stirrer (i.e., under low shear). Nanocomposite films and/or coatings with 0.1 wt % CNT exhibited improvements in electrical conductivity of 10-12 orders of magnitude compared to a pristine polymer film and a high retention of optical transparency (greater than 50%) at 500 nm. Optical microscopic examination of the nanocomposite film showed the presence of CNT bundles and agglomerates of bundles. However, the bundles were of a smaller size than those observed in nanocomposite films prepared by Methods (1)-(4). Based on a qualitative assessment, the nanocomposite film prepared via this method exhibited significantly improved dispersion relative to the nanocomposite films prepared via Methods (1)-(4). Optionally, the solution obtained by Method (5) may be filtered to remove extraneous particles or large agglomerates of CNT bundles.

Method (6) (Synthesis of the Polymer in the Presence of Pre-dispersed CNTs with Simultaneous Ultrasonic Treatment)

A combination method of preparation involving synthesis of the polymer in the presence of the CNTs while simultaneously applying ultrasonic treatment using a low power water bath operating at 40 kHz throughout the entire synthesis process was investigated. This method involved synthesis of the polymer in the presence of pre-dispersed CNTs as described in Method (5), but the reaction vessel was immersed in an ultrasonic bath throughout the entire synthesis. It should be noted that in contrast to Methods (3) and (4), which used a high power sonic horn operating at 20 kHz (100-750 Watt/cm²), the ultrasonic bath operates at a much lower level of power (less than 10 Watt/cm²) and at a higher frequency (40 kHz). Based on the observed increase in solution viscosity (indicating high molecular weight polymer formation) and microscopic analysis of the nanocomposite films, the use of the ultrasonic bath operating at 40 kHz did not cause any observable degradation of the CNTs, nor did it affect the formation of high molecular weight polymer. Nanocomposite films and/or coatings with 0.1 wt % CNT exhibited improvements in electrical conductivity of 10-12 orders of magnitude compared to a pristine polymer film and a high retention of optical transparency (i.e., greater than about 50%) at 500 nm. Optical microscopic examination of the nanocomposite film showed the presence of CNT bundles and agglomerates of bundles. However, the bundles were of a smaller size than those observed in nanocomposite films prepared by methods (1)-(4). Based on a qualitative assessment, the nanocomposite film prepared via this method exhibited significantly improved dispersion relative to the nanocomposite films prepared via Methods (1)-(4). Optionally, the solution obtained by Method (6) may be filtered to remove extraneous particles or large agglomerates of CNT bundles.

Performing synthesis of the polymers [i.e., Methods (5) and (6)] in the presence of the CNTs provided significant improvement in the dispersion of the CNTs, provided the smallest decrease in optical transmission, provided an equal or better electrical conductivity compared to a pristine polymer film and provided a stable solution. Attempts to mix a pre-synthesized high molecular weight aromatic polymer solution with a CNT dispersion was unsuccessful in achieving good dispersion and high retention of optical transmission. Methods (5) and (6) are applicable to various condensation polymers such as poly(arylene amide), polyimide and poly(arylene ether)/CNT nanocomposites as shown in FIGS. 1 and 2. FIG. 4 illustrates the preparation of polyimide and poly(amide acid)/CNT nanocomposites, wherein Ar and Ar’ can be any aromatic moiety. FIG. 2 illustrates the preparation of poly(arylene ether)/CNT nanocomposites as shown in FIGS. 1 and 2. FIG. 1 illustrates the preparation of polyimide and poly(amide acid)/CNT nanocomposites, wherein Ar and Ar’ represent any aromatic moiety, X represents a leaving group such as a halogen, nitro or other suitable group and Ar” represents any electron withdrawing group or ring system.

EXAMPLES

The following specific examples are provided for illustrative purposes and do not serve to limit the scope of the invention.

Example 1A

Preparation of 0.1 wt % CNT/polyimide nanocomposite from 1,3-bis(3-aminophenoxy)benzene (APB) and 4,4’-perfluoroisopropylidene dianhydride (6FDA) by Method (6)

FIG. 3 illustrates preparation of 0.1 wt % I.A-NT/polyimide nanocomposite from APB and 6 FDA by Method (6). Purified SWNTs obtained from Tubes@Rice as a dispersion in toluene were used as the conductive inclusions. A dilute SWNT solution, typically approximately 0.01% weight/volume (w/v) in N,N-dimethylformamide (DMF), was prepared by replacing the toluene with DMF by centrifuging and decanting several (typically three) times. Pure CNT powders could also be used, eliminating the previous step. The dilute SWNT solution was homogenized for 10 min, and sonicated for 1 hour in a ultrasonic bath operating at 40 kHz. If a higher power sonic bath is used, sonication time can be reduced depending on the power. Sonication time should be also adjusted depending on the quality of CNTs. The sonicated SWNT solution (2 mL, 0.01 g of the solid SWNT) was transferred into a 100 mL three neck round bottom flask equipped with a mechanical stirrer, nitrogen gas inlet, and drying tube outlet filled with calcium sulfate. The flask was immersed in the ultrasonic bath throughout the entire synthesis procedure. APB (3.9569 g, 1.353x10⁻² mol) was added into the flask along with 20 mL of DMF while stirring under sonication. After approximately 30 min. of stirring the SWNT and diamin mixture, 6FDA (6.0452 g, 1.360x10⁻² mol) was added along with additional 30.5 mL of DMF while stirring under sonication. The dark mixture was stirred in the sonic bath overnight, approximately 12 hours, to give a 0.1% by weight SWNT/poly(amide acid) solution. During the course of the reaction, a noticeable increase in solution viscosity was observed. The concentration of the SWNT/poly(amide acid) was 16% solids (w/w) in DMF. The SWNT/poly(amide acid) solution was treated with acetic anhydride (4.1983 g, 4.080x10⁻⁴ mol) and pyridine (3.2273 g, 1.360x10⁻² mol) to effect imidization. The resulting solution was cast onto plate glass and placed in a dry air box for 24 hours to give a tack-free film. This film was thermally treated (to remove solvent) for 1 hour each at 110, 170, 210 and 250° C. in a forced air oven. The film was removed from the glass and characterized.

Example 1B

Film was prepared in a manner identical to that described for EXAMPLE 1A, except that the SWNT concentration in the polyimide was 0.2% by weight.

Example 1C

Film was prepared in a manner identical to that described for EXAMPLE 1A, except that the SWNT concentration in the polyimide was 0.5% by weight.
Example 1D

Film was prepared in a manner identical to that described for EXAMPLE 1A, except that the SWNT concentration in the polyimide was 1.0% by weight.

Example 1E

Film was prepared in a manner to that described for EXAMPLE 1D, except that Method (1) was employed instead of Method (6).

Example 2

Preparation of 0.1 wt % LA-NT/polyimide nanocomposite from 2,6-bis(3-aminophenoxy)benzonitrile [([3-CN]APB) and ODPA by Method (6).

FIG. 4 illustrates the preparation of 0.1 wt % LA-NT/ polyimide nanocomposite from 2,6-bis(3-aminophenoxy) benzonitrile [([3-CN]APB) and ODPA by Method (6).

Prepared SWNTs obtained from Tubes@Rice as a dispersion in toluene were used as the conductive inclusions. A dilute SWNT solution, generally about 0.01% w/v in N,N-dimethylacetamide (DMAc), was prepared by replacing the toluene with DMAc by centrifuging and decanting several (typically three) times. The dilute SWNT solution was homogenized for 10 min. and sonicated for 1 hour in an ultrasonic bath operating at 40 kHz. The SWNT solution (2 mL, 0.01 g of the solid SWNT) was transferred into a 100 mL three neck round bottom flask equipped with a mechanical stirrer, nitrogen gas inlet, and drying tube outlet filled with calcium sulfate. The flask was immersed in the ultrasonic bath during the entire reaction. ([3-CN]APB, (5.0776 g, 1.60x10^-2 mol) was subsequently added to the flask along with 20 mL of DMAc and stirring under sonication overnight, approximately 12 hours, to give a 0.1 wt % SWNT/poly(amide acid) solution. During the course of the reaction, a noticeable increase in solution viscosity was observed. The concentration of the solid SWNT/poly(amide acid) solution was 16% (w/w) in DMAc. The SWNT/poly(amide acid) solution was treated with acetic anhydride (4.1983 g, 4.08x10^-2 mol) and pyridine (3.2273 g, 1.360x10^-2 mol) to effect imidization. The resulting solution was cast onto plate glass and placed in a dry air box for 24 hours to give a tack-free film. This film was thermally treated (to remove solvent) for 1 hour each at 50, 150, 200 and 250° C. in a nitrogen oven. The film was subsequently removed from the glass and characterized.

Example 3

Preparation of a 0.1% wt/wt LA-NT/polyimide nanocomposite from [2,4-bis(3-aminophenoxo)phenyl]diphenylphosphine oxide (APB-PPO) and ODPA by Method (5).

FIG. 5 illustrates preparation of a 0.1% wt/wt LA-NT/ polyimide nanocomposite APB-PPO and ODPA by Method (5).

A glass vial containing 0.060 g of nanotubes and 10 mL DMF was placed in an ultrasonic bath operating at 40 kHz for periods ranging from 16 to 24 hours. A 100 mL three neck round bottom flask equipped with a mechanical stirrer, nitrogen gas inlet, and drying tube outlet filled with calcium sulfate was charged with APB-PPO (3.6776 g, 7.467x10^-3 mole) and DMF (5.0 mL). Once the diamine dissolved, the DMF/SWNT mixture was added and the resulting mixture was stirred for 20 mins. ODPA (2.3164 g, 7.467x10^-3 mole) was added along with additional DMF (8.2 mL) to give a solution with a concentration of 20% (w/w) solids and a nanotube concentration of 0.1% wt/m. The mixture was stirred overnight at room temperature under a nitrogen atmosphere. During the course of the reaction a noticeable increase in solution viscosity was observed. The poly(amide acid) was chemically imidized by the addition of 2.31 g of acetic anhydride and 1.77 g of pyridine. The reaction mixture was stirred at room temperature overnight under a nitrogen atmosphere. The polyimide/nanomaterial mixture was precipitated in a blender containing deionized water, filtered, washed with excess water and dried in a vacuum oven at 150° C. overnight to afford a light grey, fibrous material. A solution prepared from DMF or chloroform (20% solids w/w) was cast onto plate glass and allowed to dry to a tack-free state in a dust-free chamber. The film on the glass plate was placed in a forced air oven for 1 hour each at 100, 150, 175 and 225° C. to remove solvent. The film was subsequently removed from the glass and characterized.

Example 4

Preparation of a 0.2% wt/wt LA-NT/polyimide nanocomposite from APB-PPO and ODPA via Method (5).

FIG. 6 illustrates preparation of a 0.2% wt/wt LA-NT/ polyimide nanocomposite from APB-PPO and ODPA via method (5).

A glass vial containing 0.0120 g of LA-NT nanotubes and 10 mL of DMF was placed in an ultrasonic bath operating at 40 kHz for periods ranging from 16 to 24 hours. A 100 mL three neck round bottom flask equipped with a mechanical stirrer, nitrogen gas inlet, and drying tube outlet filled with calcium sulfate was charged with APB-PPO (3.6776 g, 7.467x10^-3 mole) and DMF (5.0 mL). Once the diamine dissolved, the DMF/SWNT mixture was added and the resulting mixture was stirred for 20 min. ODPA (2.3164 g, 7.467x10^-3 mole) was added along with additional DMF (8.2 mL) to give a solution with a concentration of 20% (w/w) solids and a nanotube concentration of 0.2% wt/wt. The mixture was stirred overnight at room temperature under a nitrogen atmosphere. The poly(amide acid) was chemically imidized by the addition of 2.31 g of acetic anhydride and 1.77 g of pyridine. The reaction mixture was stirred at room temperature overnight, approximately 12 hours, under a nitrogen atmosphere. During the course of the reaction a noticeable increase in solution viscosity was observed. The polyimide/SWNT mixture was precipitated in a blender containing deionized water, filtered, washed with excess water and dried in a vacuum oven at 150° C. overnight to afford a light grey, fibrous material. A solution prepared from DMF or chloroform (20% solids w/w) was cast onto plate glass and allowed to dry to a tack-free state in a dust-free chamber. The film on the glass plate was placed in a forced air oven for 1 hour each at 100, 150, 175 and 225° C. to remove solvent. The film was subsequently removed from the glass and characterized.
Preparation of a 0.1% wt/wt CVD-NT-1/polyimide nanocomposite from APB-PPO and ODPA by Method (5)

FIG. 7 illustrates the preparation of a 0.1% wt/wt CVD-NT-1/polyimide nanocomposite from APB-PPO and ODPA by Method (5).

A glass vial containing 0.0060 g of CVD-NT-1 nanotubes and 10 mL of DMF was placed in an ultrasonic bath at 40 kHz for periods ranging from 16 to 24 hours. A 100 mL three neck round bottom flask equipped with a mechanical stirrer, nitrogen gas inlet, and drying tube filled with calcium sulfate was charged with APB-PPO (3.6776 g, \(7.467 \times 10^{-3}\) mole) and DMF (5.0 mL). Once the diamine dissolved, the DMF/CNT mixture was added and the resulting mixture was stirred for 20 min. ODPA (2.3164 g, \(7.467 \times 10^{-3}\) mole) was added along with additional DMF (8.2 mL) to give a solution with a concentration of 20% (w/w) solids and a nanotube concentration of 0.1% wt/wt. The mixture was stirred overnight at room temperature under a nitrogen atmosphere. During the course of the reaction, a noticeable increase in solution viscosity was observed. The poly(amide acid) was chemically imidized by the addition of 2.31 g of acetic anhydride and 1.77 g of pyridine. The reaction mixture was stirred at room temperature overnight under a nitrogen atmosphere. The polyimide/CNT solution was precipitated in a blender containing deionized water, filtered, washed with excess water and dried in a vacuum oven at 150° C. overnight to afford a light gray, fibrous material. A solution prepared from DMF or chloroform (20% solids w/w) was cast onto plate glass and allowed to dry to a tack-free state in a dust-free chamber. The film on the glass plate was placed in a forced air oven for 1 hour each at 100, 150, 175 and 225° C. to remove solvent. The film was subsequently removed from the glass and characterized.

Example 6

Preparation of a 0.2% wt/wt CVD-NT-1/polyimide nanocomposite from APB-PPO and ODPA by Method (5)

FIG. 8 illustrates the preparation of a 0.2% wt/wt CVD-NT-1/polyimide nanocomposite from APB-PPO and ODPA by Method (5).

A glass vial containing 0.0120 g of CVD-NT-1 nanotubes and 10 mL of DMF was placed in an ultrasonic bath operating at 40 kHz for periods ranging from 16 to 24 hours. A 100 mL three neck round bottom flask equipped with a mechanical stirrer, nitrogen gas inlet, and drying tube filled with calcium sulfate was charged with APB-PPO (3.6776 g, \(7.467 \times 10^{-3}\) mole) and DMF (5.0 mL). Once the diamine dissolved, the DMF/CNT mixture was added and the resulting mixture was stirred for 20 min. ODPA (2.3164 g, \(7.467 \times 10^{-3}\) mole) was added along with additional DMF (8.2 mL) to give a solution with a concentration of 20% (w/w) solids and a nanotube concentration of 0.2% wt/wt. The mixture was stirred overnight, approximately 12 hours, at room temperature under a nitrogen atmosphere. During the course of the reaction, a noticeable increase in solution viscosity was observed. The poly(amide acid) was chemically imidized by the addition of 2.31 g of acetic anhydride and 1.77 g of pyridine. The reaction mixture was stirred at room temperature overnight under a nitrogen atmosphere. The polyimide/CNT solution was precipitated in a blender containing deionized water, filtered, washed with excess water and dried in a vacuum oven at 150° C. overnight to afford a light gray, fibrous material. A solution prepared from DMF or chloroform (20% solids w/w) was cast onto plate glass and allowed to dry to a tack-free state in a dust-free chamber. The film on the glass plate was placed in a forced air oven for one hour each at 100, 150, 175 and 225° C. to remove solvent. The film was subsequently removed from the glass and characterized.

Example 7

Preparation of a 0.1% wt/wt CVD-NT-2/polyimide nanocomposite from APB-PPO and ODPA by Method (5)

FIG. 9 illustrates the preparation of a 0.1% wt/wt CVD-NT-2/polyimide nanocomposite from APB-PPO and ODPA by Method (5).

A glass vial containing 0.0060 g of CVD-NT-2 nanotubes and 10 mL of DMF was placed in an ultrasonic bath operating at 40 kHz for periods ranging from 16 to 24 hours. A 100 mL three neck round bottom flask equipped with a mechanical stirrer, nitrogen gas inlet, and drying tube filled with calcium sulfate was charged with APB-PPO (3.6776 g, \(7.467 \times 10^{-3}\) mole) and DMF (5.0 mL). Once the diamine dissolved, the DMF/CNT mixture was added and the resulting mixture was stirred for 20 minutes. ODPA (2.3164 g, \(7.467 \times 10^{-3}\) mole) was added along with additional DMF (8.2 mL) to give a solution with a concentration of 20% (w/w) solids and a nanotube concentration of 0.1% wt/wt. The mixture was stirred overnight, approximately 12 hours, at room temperature under a nitrogen atmosphere. During the course of the reaction, a noticeable increase in solution viscosity was observed. The poly(amide acid) was chemically imidized by the addition of 2.31 g of acetic anhydride and 1.77 g of pyridine. The reaction mixture was stirred at room temperature overnight under a nitrogen atmosphere. The polyimide/CNT solution was precipitated in a blender containing deionized water, filtered, washed with excess water and dried in a vacuum oven at 150° C. overnight to afford a light gray, fibrous material. A solution prepared from DMF or chloroform (20% solids w/w) was cast onto plate glass and allowed to dry to a tack-free state in a dust-free chamber. The film on the glass plate was placed in a forced air oven for one hour each at 100, 150, 175 and 225° C. to remove solvent. The film was subsequently removed from the glass and characterized.

Example 8

Preparation of a 0.2% wt/wt CVD-NT-2/polyimide nanocomposite from APB-PPO and ODPA by Method (5)

FIG. 10 illustrates preparation of a 0.2% wt/wt CVD-NT-2/polyimide nanocomposite from APB-PPO and ODPA by Method (5).

A glass vial containing 0.0120 g of CVD-NT-2 nanotubes and 10 mL of DMF was placed in an ultrasonic bath operating at 40 kHz for periods ranging from 16 to 24 hours. A 100 mL three neck round bottom flask equipped with a mechanical stirrer, nitrogen gas inlet, and drying tube filled with calcium sulfate was charged with APB-PPO (3.6776 g, \(7.467 \times 10^{-3}\) mole) and DMF (5.0 mL). Once the diamine dissolved, the DMF/nanomaterial mixture was added and the resulting mixture was stirred for 20 minutes. ODPA (2.3164 g, \(7.467 \times 10^{-3}\) mole) was added along with additional DMF (8.2 mL) to give a solution with a concentration of 20% (w/w) solids and a nanotube concentration of 0.2% wt/wt. The mixture was stirred overnight, approximately 12 hours, at room temperature under a nitrogen atmosphere. During the course of the reaction, a noticeable increase in solution viscosity was observed. The poly(amide acid) was chemically imidized by the addition of 2.31 g of acetic anhydride and 1.77 g of pyridine. The reaction mixture was stirred at room temperature overnight under a nitrogen atmosphere. The polyimide/CNT solution was precipitated in a blender containing deionized water, filtered, washed with excess water and dried in a vacuum oven at 150° C. overnight to afford a light gray, fibrous material. A solution prepared from DMF or chloroform (20% solids w/w) was cast onto plate glass and allowed to dry to a tack-free state in a dust-free chamber. The film on the glass plate was placed in a forced air oven for one hour each at 100, 150, 175 and 225° C. to remove solvent. The film was subsequently removed from the glass and characterized.
Example 9
Preparation of a 0.1% wt/wt. LA-NT/poly(arylene ether)/SWNT nanocomposite by Method (5)

FIG. 11 illustrates the preparation of a 0.1% wt/wt LA-NT/ poly(arylene ether)/SWNT nanocomposite by Method (5).

Example 10
Preparation of 0.1 wt% CNT/PMMA nanocomposite from methyl methacrylate (MMA) monomers by Method (6)

Purified SWNTs obtained from Tubes@Rice were used as the conductive inclusions. A dilute SWNT solution, typically approximately 0.01% weight/weight (w/w) in N,N-dimethylformamide (DMF), was prepared. The dilute SWNT solution was homogenized for 10 min and sonicated for one hour in an ultrasonic bath operating at 40 kHz. The sonicated SWNT solution (2 mL, 0.01 g of the solid SWNT) was transferred into a 100 mL three neck round bottom flask equipped with a mechanical stirrer, nitrogen gas inlet, and a cooling trap. A solution prepared from MMA (20% solids w/w) was added onto plate glass and allowed to dry to a tack-free state in a dust-free chamber. The film on the glass plate was placed in a forced air oven for one hour each at 100, 150, 175 and 225 °C. to remove solvent. The film was subsequently removed from the glass and characterized.

Preparation of a 0.1% wt/wt. LA-NT/poly(arylene ether)/SWNT nanocomposite by Method (5)

The above examples are provided for illustrative purposes. In addition to the specific condensation and addition polymers described herein, other addition and condensation polymers may be used, including polymides, polysters, polycarbonates, vinyl polymers, polyethylene, polycrilonitrile, poly(vinyl chloride), polystyrene, poly(vinyl acetate), poly(tetrafluoroethylene), polyisprene, polyurethane, and poly(methyl methacrylate)/polystyrene copolymer.

Characterization

Differential scanning calorimetry (DSC) was conducted on a Shimadzu DSC-50 thermal analyzer. The glass transition temperature (Tg) was taken as the inflection point of the ΔT versus temperature curve at a heating rate of 10 °C/min on thin film samples. UV/VIS spectra were obtained on thin films using a Perkin-Elmer Lambda 900 UV/VIS/NIR spectrophotometer. Thin-film tensile properties were determined according to ASTM D882 using four specimens per test condition. Thermogravimetric analysis (TGA) was performed on a Seiko Model 200/220 instrument on film samples at a heating rate of 2.5 °C/min in air and/or nitrogen at a flow rate of 15 cm³ min⁻¹. Conductivity measurements were performed according to ASTM D257 using a Keithley 6517 Electrometer. Homogenization was carried out using PowerGen Model 35 or a PowerGen Model 700 homogenizer at speeds ranging from 5,000 to 30,000 rpm. Optionally, a fluidizer, such as a M-10Y High Pressure Microfluidizer from WIC Corp. (Newton, Mass.) could be used. Solar absorptivities were measured on an Aztek Model LPSR-300 spectrorreflectometer with measurements taken between 250 to 2800 nm with a vapor deposited aluminum on Kapton® as a reflective reference. An Aztek Temp 2000A Infrared reflectometer was used to measure the thermal emissivity. Ultrasonication was carried out using an Ultrasonic 57x ultrasonicator water bath operating at 40 kHz or with an ultrasonic horn (VCX-750, Sonics and Materials, Inc.) equipped with a 13 millimeter probe. Purified, laser ablated single wall carbon nanotubes (LA-NT) were used as received from Tubes@Rice, Rice University, Houston, Tex. Chemical vapor deposition multi-wall carbon nanotubes (CVD-NT) were used as received from Nanolab, Inc., Watertown, Mass. Optical microscopy was performed on an Olympus BH-2 microscope. Elemental analysis was performed by Desert Analytics, Tucson, Ariz.
The nanocomposite films were characterized for optical, electrical and thermal properties. CNTs were subjected to elemental analysis prior to use. The results are summarized in Table 1. Characterization of the nanocomposite films described in EXAMPLES 1A-1E are presented in Tables 2 and 3. All of these samples were prepared using LA purified SWNTs obtained from Tubes@Rice.

<table>
<thead>
<tr>
<th>CNT</th>
<th>Carbon, %</th>
<th>Hydrogen, %</th>
<th>Iron, %</th>
<th>Nickel, %</th>
<th>Cobalt, %</th>
</tr>
</thead>
<tbody>
<tr>
<td>LA-NT</td>
<td>78.2</td>
<td>0.94</td>
<td>0.06</td>
<td>1.45</td>
<td>1.54</td>
</tr>
<tr>
<td>CVD-NT-1</td>
<td>96.0</td>
<td>&lt;0.05</td>
<td>1.0</td>
<td>0.002</td>
<td>&lt;0.001</td>
</tr>
<tr>
<td>CVD-NT-2</td>
<td>97.0</td>
<td>&lt;0.05</td>
<td>1.5</td>
<td>0.002</td>
<td>&lt;0.001</td>
</tr>
</tbody>
</table>

Table 4 presents the dynamic mechanical data of the nanocomposite films. Table 5 lists the types, sources and approximate dimensions of the nanotubes used in the preparation of nanocomposite films described in EXAMPLES 3-8.

<table>
<thead>
<tr>
<th>Sample Film</th>
<th>SWNT Loading, Weight %</th>
<th>Temp. of 5% Weight Loss, °C</th>
<th>Solar absorptivity (α)</th>
<th>Thermal emissivity (ε)</th>
</tr>
</thead>
<tbody>
<tr>
<td>APB/6FDA</td>
<td>6</td>
<td>444</td>
<td>0.068</td>
<td>0.525</td>
</tr>
<tr>
<td>EXAMPLE 1A</td>
<td>0.1</td>
<td>461</td>
<td>0.268</td>
<td>0.578</td>
</tr>
<tr>
<td>EXAMPLE 1B</td>
<td>0.2</td>
<td>474</td>
<td>0.398</td>
<td>0.614</td>
</tr>
<tr>
<td>EXAMPLE 1C</td>
<td>0.5</td>
<td>481</td>
<td>0.362</td>
<td>0.620</td>
</tr>
<tr>
<td>EXAMPLE 1D</td>
<td>1.0</td>
<td>479</td>
<td>0.478</td>
<td>0.652</td>
</tr>
</tbody>
</table>

Another series of 0.1 and 0.2 wt % nanocomposite films were prepared from the polyimide derived from APB-PPO and ODPA and three different types of CNTs. Method (5) was used for the preparation of the nanocomposite films described in EXAMPLES 3-9. The nanotubes differed in their method of preparation (either LA or CVD) as well as the average lengths and diameters of the tubes. In addition, LA-NT are single wall carbon nanotubes (SWNTs) and CVD-NT-1 and CVD-NT-2 are multi-wall carbon nanotubes (MWNTs). Table 6 lists physical properties of the nanocomposite films, such as Φ and thin film mechanical properties. The Tg ranged from 187 to 212° C. The films exhibited room temperature tensile strengths and moduli from 77 to 99 MPa and 2.8 to 3.3 GPa, respectively. The elongations at break ranged from 3.1 to 4.9%. These values are comparable to other aromatic polyimides. The polyimide/CNT nanocomposite films exhibited reductions in Tg of 5-25° C., comparable tensile strengths (except for EXAMPLE 6), increased tensile moduli and comparable or slightly lower elongations to break.

Table 5 lists the types, sources and approximate dimensions of the nanotubes used in the preparation of nanocomposite films described in EXAMPLES 3-8.

<table>
<thead>
<tr>
<th>Nanotube Designations, Source and Approximate Dimensions</th>
</tr>
</thead>
<tbody>
<tr>
<td>Nanotube ID (Type)</td>
</tr>
<tr>
<td>---------------------</td>
</tr>
<tr>
<td>LA-NT (SWNT)</td>
</tr>
<tr>
<td>CVD-NT-1 (MWNT)</td>
</tr>
<tr>
<td>CVD-NT-2 (MWNT)</td>
</tr>
</tbody>
</table>

The tan δ peak decreased and shifted up 10° C. with SWNT incorporation at 1.0 vol % as seen in Table 4, which suggests that CNT reinforcement made the nanocomposite more elastic and thermally more stable by increasing the glass transition temperature.

Table 3 lists the types, sources and approximate dimensions of the nanotubes used in the preparation of nanocomposite films described in EXAMPLES 3-8.

### Table 3

<table>
<thead>
<tr>
<th>Sample Film</th>
<th>SWNT Loading, Weight %</th>
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<td>EXAMPLE 1C</td>
<td>0.5</td>
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<td>0.362</td>
<td>0.620</td>
</tr>
<tr>
<td>EXAMPLE 1D</td>
<td>1.0</td>
<td>479</td>
<td>0.478</td>
<td>0.652</td>
</tr>
</tbody>
</table>

### Table 4

<table>
<thead>
<tr>
<th>Sample Film</th>
<th>Tan δ Max, °C</th>
<th>Storage modulus (GPa)</th>
</tr>
</thead>
<tbody>
<tr>
<td>APB/6FDA</td>
<td>214</td>
<td>8.5 x 10^9</td>
</tr>
<tr>
<td>EXAMPLE 1A</td>
<td>213</td>
<td>9.2 x 10^9</td>
</tr>
<tr>
<td>EXAMPLE 1C</td>
<td>214</td>
<td>1.2 x 10^9</td>
</tr>
<tr>
<td>EXAMPLE 1D</td>
<td>224</td>
<td>1.4 x 10^9</td>
</tr>
</tbody>
</table>

### Table 5

<table>
<thead>
<tr>
<th>Nanotube Designations, Source and Approximate Dimensions</th>
</tr>
</thead>
<tbody>
<tr>
<td>Nanotube ID (Type)</td>
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<tr>
<td>---------------------</td>
</tr>
<tr>
<td>LA-NT (SWNT)</td>
</tr>
<tr>
<td>CVD-NT-1 (MWNT)</td>
</tr>
<tr>
<td>CVD-NT-2 (MWNT)</td>
</tr>
</tbody>
</table>

Table 6 lists physical properties of the nanocomposite films, such as Φ and thin film mechanical properties. The Tg ranged from 187 to 212° C. The films exhibited room temperature tensile strengths and moduli from 77 to 99 MPa and 2.8 to 3.3 GPa, respectively. The elongations at break ranged from 3.1 to 4.9%. These values are comparable to other aromatic polyimides. The polyimide/CNT nanocomposite films exhibited reductions in Tg of 5-25° C., comparable tensile strengths (except for EXAMPLE 6), increased tensile moduli and comparable or slightly lower elongations to break.
Imidized thin film samples were measured for optical transparency using UV/VIS spectroscopy. The results are presented in Table 7. The retention of optical transparency at 500 nm ranged from 52 to 89%. It is well known that for these polyimide films, the optical transmission is dependent upon film thickness such that increasing film thickness results in a decrease in optical transmission. As shown in Table 7, the films thicknesses of the nanocomposite films were comparable or slightly greater than that of the control. Thus it is reasonable to compare the results directly without normalization. "High", "moderate", and "poor" retention of optical transparency are defined herein to mean greater than 50%, 35% to 50%, and less than 35%, respectively.

**TABLE 7**

<table>
<thead>
<tr>
<th>Sample Film, (CNT conc., wt %)</th>
<th>Relative CNT Loading, %</th>
<th>Relative Transparency @ 500 nm, %</th>
<th>Relative Retention of Optical Transparency, %</th>
<th>Film Thickness, µm</th>
</tr>
</thead>
<tbody>
<tr>
<td>APB-PPO/ODPA</td>
<td>0</td>
<td>85</td>
<td></td>
<td>25</td>
</tr>
<tr>
<td>EXAMPLE 3</td>
<td>0.1</td>
<td>76</td>
<td>89</td>
<td>32</td>
</tr>
<tr>
<td>EXAMPLE 5</td>
<td>0.1</td>
<td>66</td>
<td>78</td>
<td>32</td>
</tr>
<tr>
<td>EXAMPLE 7</td>
<td>0.1</td>
<td>48</td>
<td>56</td>
<td>32</td>
</tr>
<tr>
<td>EXAMPLE 4</td>
<td>0.2</td>
<td>75</td>
<td>88</td>
<td>27</td>
</tr>
<tr>
<td>EXAMPLE 6</td>
<td>0.2</td>
<td>44</td>
<td>52</td>
<td>25</td>
</tr>
</tbody>
</table>

Thermal emissivity (ε) and solar absorptivity (α) measurements are shown in Table 8. In general, the addition of CNTs to the polyimide material increased both ε and α. The solar absorptivity increased depending upon CNT type, for example the samples with the laser ablated nanotubes (SWNTs, EXAMPLES 3 and 4) exhibited the lowest increase while the chemical vapor deposited nanotubes (MWNTs, EXAMPLES 5 and 6) exhibited the largest increase.

Although the present invention has been described in detail, it should be understood that various changes, substitutions, and alterations may be readily ascertainable by those skilled in the art and may be made herein without departing from the spirit and scope of the present invention as defined by the following claims.

What is claimed as new and desired to be secured by Letters Patent of the United States is:

1. An electrically conductive, optically transparent poly(amide acid)/carbon nanotube nanocomposite, comprising a carbon nanotube and a poly(amide acid), wherein the poly(amide acid) is

\[
\text{Ar} - \text{N} = \text{N} - \text{Ar}'
\]

wherein Ar is selected from the group consisting of

\[
\begin{align*}
\text{CN} & \quad \text{O} & \quad \text{O} & \quad \text{O} & \quad \text{O} \\
\text{O} & \quad \text{O} & \quad \text{O} & \quad \text{O} & \quad \text{O} \\
\end{align*}
\]

wherein Ar' is selected from the group consisting of

\[
\begin{align*}
\text{CF}_3 & \quad \text{CF}_3 & \quad \text{CF}_3 \\
\end{align*}
\]
wherein the number average molecular weight of the poly
(amide acid) can range from approximately 700 g/mole
to approximately 100,000 g/mole; and
wherein the poly(amide acid) is selected from the group
consisting of endcapped and unendcapped.

2. An electrically conductive, optically transparent poly-
imide/carbon nanotube nanocomposite comprising a carbon
nanotube and a polyimide, wherein the polyimide is

wherein \( \text{Ar} \) is selected from the group consisting of

3. An electrically conductive, optically transparent poly
(arylene ether)/carbon nanotube nanocomposite comprising
a carbon nanotube and a poly(arylene ether), wherein the
poly(arylene ether) is

wherein \( \text{Ar}'' \) is

4. A nanocomposite product prepared from the nanocom-
posite of claim 1, wherein said product is in a form selected
from the group consisting of a film, fiber, foam, coating,
adhesive, molding and paste.

5. A nanocomposite product prepared from the nanocom-
posite of claim 2, wherein said product is in a form selected
from the group consisting of a film, fiber, foam, coating,
adhesive, molding and paste.

6. A nanocomposite product prepared from the nanocom-
posite of claim 3, wherein said product is in a form selected
from the group consisting of a film, fiber, foam, coating,
adhesive, molding and paste.