There is provided in one embodiment of the invention an electrolyte for use in a lithium ion electrochemical cell. The electrolyte comprises a mixture of an ethylene carbonate (EC), an ethyl methyl carbonate (EMC), an ester cosolvent, and a lithium salt. The ester cosolvent comprises methyl propionate (MP), ethyl propionate (EP), methyl butyrate (MB), ethyl butyrate (EB), propyl butyrate (PB), or butyl butyrate (BB). The electrochemical cell operates in a temperature range of from about -60 degrees Celsius to about 60 degrees Celsius. In another embodiment there is provided a lithium ion electrochemical cell using the electrolyte of the invention.
References Cited

U.S. PATENT DOCUMENTS

<table>
<thead>
<tr>
<th>Patent Number</th>
<th>Date</th>
<th>Inventors</th>
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<tbody>
<tr>
<td>5,478,673 A</td>
<td>12/1995</td>
<td>Funatsu</td>
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<tr>
<td>6,489,055 B1</td>
<td>12/2002</td>
<td>Ichishahi et al.</td>
</tr>
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<td>6,492,064 B1</td>
<td>12/2002</td>
<td>Smart et al.</td>
</tr>
<tr>
<td>2002/009651 A1</td>
<td>1/2002</td>
<td>Barker et al.</td>
</tr>
<tr>
<td>2008/030546 A1</td>
<td>12/2008</td>
<td>Smart et al.</td>
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<td>2012/000756 A1</td>
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FOREIGN PATENT DOCUMENTS

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<tr>
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<tr>
<td>JP 9-17446</td>
<td>1/1997</td>
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<tr>
<td>WO 2008/109802</td>
<td>9/2008</td>
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OTHER PUBLICATIONS

Colbow et al. (May 16, 1989) “Structure and Electrochemistry of the Spinel Oxides Li2Ti3O7 and Li7Ti5O12,” J Power Sources 26(3-4):397-402.
Smart et al. (Oct. 13, 2010) “Improved Wide Operating Temperature Range of MCMB-Li1.1(Co0.4Ni0.4Mn0.2)O2 Cells with Methyl Butyrate-Based Electrolytes,” Abstract 580, 218th Meeting of the Electrochemical Society, Las Vegas, NV.

* cited by examiner
Fig. 2

Graph showing voltage (V) vs. percentage of room temperature capacity (°C) for MCMB carbon-LiNiO₂ cells. The graph includes multiple lines representing different solutions and temperatures.

50 mA discharge current to 2.00 V
TEMP = 40°C
Fig. 6

MCMB CARBON-LiNiCoO₂ CELLS

5 mA DISCHARGE CURRENT TO 2.00 V.
TEMP = -60° C

VOLTAGE (V)

PERCENTAGE OF ROOM TEMPERATURE CAPACITY (%)
MCMB CARBON-LiNiO₂ CELLS

10 mA DISCHARGE CURRENT TO 2.00 V
TEMP = -60°C

PERCENTAGE OF ROOM TEMPERATURE CAPACITY

VOLTAGE (V)

Fig. 7
Fig. 8

MCMB CARBON-LiNiCoO₂ CELLS
25 mA DISCHARGE CURRENT TO 2.00 V
TEMP = 40° C

PERCENTAGE OF ROOM TEMPERATURE CAPACITY (%)
Fig. 9

MCMB CARBON-LiNi$_x$Co$_{1-x}$O$_2$ CELLS
LITHIUM REFERENCE ELECTRODE
TEMPERATURE = 23°C

CATHODE POTENTIAL (V vs. Li/Li$^+$)

LiNi$_x$Co$_{1-x}$O$_2$ CATHODES

<table>
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<tr>
<th>Current (Amps)</th>
<th>0.001</th>
<th>0.01</th>
<th>0.1</th>
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<td></td>
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<tr>
<td>1.0 M LiPF$_6$ EC+EMC (20:80 v/v %)</td>
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<td>1.0 M LiPF$_6$ EC+EMC+EP (20:60:20 v/v/v %)</td>
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<tr>
<td>1.0 M LiPF$_6$ EC+EMC+MB (20:60:20 v/v/v %)</td>
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<tr>
<td>1.0 M LiPF$_6$ EC+EMC+EB (20:60:20 v/v/v %)</td>
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<td>1.0 M LiPF$_6$ EC+EMC+PB (20:60:20 v/v/v %)</td>
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Fig. 11

MCMB CARBON-LiNi$_x$Co$_{1-x}$O$_2$ CELLS

LITHIUM REFERENCE ELECTRODE TEMPERATURE = 40°C

CATHODE POTENTIAL (V vs. Li$^+/Li$)

LINI$_x$Co$_{1-x}$O$_2$ CATHODES

CURRENT (AMPS)

- 1.0 M LiPF$_6$ EC+DEC+DMC (1:1:1 v/v)
- 1.0 M LiPF$_6$ EC+EMC+MP (20:60:20 v/v/v)
- 1.0 M LiPF$_6$ EC+EMC+MB (20:60:20 v/v/v)
- 1.0 M LiPF$_6$ EC+EMC+EB (20:60:20 v/v/v)
- 1.0 M LiPF$_6$ EC+EMC+PP (20:60:20 v/v/v)
**Fig. 14**

**MCMB CARBON-LiNi\textsubscript{x}Co\textsubscript{1-x}O\textsubscript{2} CELLS**

LITHIUM REFERENCE ELECTRODE

TEMPERATURE = -60°C

**ANODE POTENTIAL**

(mV vs. Li+/Li)

**CURRENT**

(AMPS)

- ■ 1.0 M LiPF\textsubscript{6} EC+EMC (20:80 v/v %)
- ○ 1.0 M LiPF\textsubscript{6} EC+DEC+DMC (1:1:1 v/v %)
- ○ 1.0 M LiPF\textsubscript{6} EC+EMC+MP (20:60:20 v/v/v %)
- ● 1.0 M LiPF\textsubscript{6} EC+EMC+EP (20:60:20 v/v/v %)
- ★ 1.0 M LiPF\textsubscript{6} EC+EMC+MB (20:60:20 v/v/v %)
- □ 1.0 M LiPF\textsubscript{6} EC+EMC+EB (20:60:20 v/v/v %)
- ▲ 1.0 M LiPF\textsubscript{6} EC+EMC+PB (20:60:20 v/v/v %)
- ▼ 1.0 M LiPF\textsubscript{6} EC+EMC+BB (20:60:20 v/v/v %)
Fig. 24

LITHIUM METAL REFERENCE ELECTRODE

MCMB CARBON-LiNi,Co,Fe₂O₄ ELECTRODE

TEMPERATURE = 40°C

MCMB ANODE MEASUREMENTS

Z' (Ohms)

Z'' (Ohms)

L-ION EXPERIMENTAL THREE ELECTRODE CELL

0.1 M LiPF₆ EC + EMC (20:80 V/V %)
0.1 M LiPF₆ EC + EMC-MP (20:80:20 V/V %)
0.1 M LiPF₆ EC + EMC + EP (20:80:20 V/V %)
0.1 M LiPF₆ EC + EMC + MB (20:80:20 V/V %)
0.1 M LiPF₆ EC + EMC + EB (20:80:20 V/V %)
0.1 M LiPF₆ EC + EMC + PB (20:80:20 V/V %)
0.1 M LiPF₆ EC + EMC + BB (20:80:20 V/V %)
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<th>CELL NUMBER</th>
<th>CELL ID</th>
<th>CELL WEIGHT (GRAMS)</th>
<th>CELL WEIGHT (g)</th>
<th>INITIAL VOLTAGE</th>
<th>INITIAL CAPACITY (Ah)</th>
<th>INITIAL Wt/hr</th>
<th>CALCULATED IMPEDANCE (mOhms at 100% SOC)</th>
<th>CALCULATED IMPEDANCE (mOhms at 75% SOC)</th>
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<td>7-3-242</td>
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<td>0.2434</td>
<td>3.554</td>
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<td>118.68</td>
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<td>0.2437</td>
<td>3.555</td>
<td>7.9548</td>
<td>28.714</td>
<td>117.84</td>
<td>16.49</td>
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<td>28.211</td>
<td>116.31</td>
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<td>7.9002</td>
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<td>18.66</td>
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<td>118.60</td>
<td>13.85</td>
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<tr>
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**AVERAGE** 243.16 0.24 3.56 7.97 28.80 118.43 16.03 15.07
Fig. 31

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<th>Lithium/Yardney 7 Ah Li-Ion Cells</th>
<th>MCMB Carbon-LiNiCoO2</th>
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<tr>
<td>C/5 DISCHARGE RATE</td>
<td>TEMPERATURE = -50°C</td>
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<tr>
<td>1400 AMP CHARGE CURRENT (C/5) TO 4.1 V</td>
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</tr>
<tr>
<td>TAPER CUT-OFF AT 0.025 A (C/20)</td>
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</tr>
<tr>
<td>CELL CHARGED AT RT PRIOR TO LT DISCHARGE</td>
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</tr>
<tr>
<td>1400 A DISCHARGE CURRENT</td>
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</table>

Discharge Capacity (Ah)

- CELL LW358: 1.0 LiPF6 EC+EMC+MP (20:60:20 w/w %)
- CELL LW266: 1.0 LiPF6 EC+EMC+EB (20:60:20 w/w %)
- CELL LW286: 1.0 LiPF6 EC+EMC+EP (20:60:20 w/w %)
- CELL LW226: 1.0 LiPF6 EC+DEC+DMC+EMC (1:1:1:3 w/v %)
- CELL LW242: 1.0 LiPF6 EC+DEC+DMC (1:1:1 w/v %)

2003 MIR ROVER BATTERY BASELINE
Fig. 32

Lithium/Yardney 7 Ah Li-Ion Cells
MCMB Carbon-LNiCoO2

1,400 AMP CHARGE CURRENT (C/5) TO 4.1 V
TAPER CUT-OFF AT 0.025 A (~ C/280)
CELL CHARGED AT RT PRIOR TO LT DISCHARGE

1,400 A DISCHARGE CURRENT
C/5 DISCHARGE RATE
TEMPERATURE = -50°C

CELL VOLTAGE (V)

DISCHARGE ENERGY (Wh/Kg)

- CELL LW358 - 1.0 LiPF6 EC+EMC+MP (20:60:20 v/v/v %)
- CELL LW361 - 1.0 LiPF6 EC+EMC+EP (20:60:20 v/v/v %)
- CELL LW242 - 1.0 LiPF6 EC+DEC+DMC+EMC (1:1:1:1 v/v %)
- CELL LW365 - 1.0 LiPF6 EC+EMC+EB (20:60:20 v/v/v %)
- CELL LW256 - 1.0 LiPF6 EC+DEC+DMC+EMC (1:1:1:1 v/v %)

2003 MER
ROVER BATTERY
BASELINE
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<tr>
<th>ELECTROLYTE TYPE</th>
<th>CHARGE CAPACITY (Ah) 1st CYCLE</th>
<th>DISCHARGE CAPACITY (Ah) 1st CYCLE</th>
<th>IRREVERSIBLE CAPACITY EFFICIENCY (1st CYCLE)</th>
<th>COULOMBIC CAPACITY (Ah) 5th CYCLE</th>
<th>REVERSIBLE CAPACITY (Ah) 5th CYCLE</th>
<th>CUMULATIVE IRREVERSIBLE CAPACITY (1st-5th CYCLE)</th>
<th>COULOMBIC EFFICIENCY (5th CYCLE)</th>
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<td>0.4788</td>
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<td>85.75</td>
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<tr>
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<td>0.076</td>
<td>83.58</td>
<td>0.3867</td>
<td>0.3737</td>
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<tr>
<td>1.0 M LiPF6 EC+EMC (20:80 v/v %)</td>
<td>0.4826</td>
<td>0.4132</td>
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<td>85.62</td>
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<td>1.0 M LiPF6 EC+EMC</td>
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LITHIUM ION ELECTROLYTES AND LITHIUM ION CELLS WITH GOOD LOW TEMPERATURE PERFORMANCE

CROSS-REFERENCE TO RELATED APPLICATION

This application claims priority of U.S. Provisional Patent Application Ser. No. 61/123,441, filed Apr. 8, 2008, which is incorporated herein by reference in its entirety.

GOVERNMENT INTERESTS

The invention described herein was made in the performance of work under a NASA contract, and is subject to the provisions of Public Law 96-517 (35 U.S.C. 202) in which the Contractor has elected to retain title.

BACKGROUND OF THE INVENTION

a. Field of the Invention

The invention relates to electrolytes and organic solvents for electrochemical cells. In particular, the invention relates to lithium ion electrolytes and organic solvents for lithium ion cells.

b. Background Art

Lithium ion cells typically include a carbon (e.g., coke or graphite) anode intercalated with lithium ions to form Li_x C; an electrolyte consisting of a lithium salt dissolved in one or more organic solvents; and a cathode made of an electrochemically active material, typically an insertion compound, such as LiCoO_2. During cell discharge, lithium ions pass from the carbon anode, through the electrolyte to the cathode, where the ions are taken up with the simultaneous release of electrical energy. During cell recharge, lithium ions are transferred back to the anode, where they reintercalate into the carbon matrix.

Lithium ion rechargeable batteries have the demonstrated characteristics of high energy density, high voltage, and excellent cycle life. Known state-of-the-art lithium ion rechargeable batteries and systems have been demonstrated to operate over a wide range of temperatures (e.g., −30° C. to +40° C.). However, the performance of such known lithium ion rechargeable batteries and systems is limited at temperatures below −30° C., making them unsuitable for many terrestrial and extra-terrestrial applications. Many scheduled NASA missions demand good low temperature battery performance without sacrificing such properties as light weight, high specific energy, long cycle life, and moderate cost. Moreover, such scheduled NASA missions require rechargeable batteries that can operate at low temperatures to satisfy the requirements of various applications, such as landers, rovers, and penetrators. For example, the Mars Exploration Program requires rechargeable batteries capable of delivering several hundred cycles with high specific energy, and the ability to operate over a broad range of temperatures, including the extremely low temperatures on and beneath the surface of Mars. Mars rovers and landers require batteries that can operate at temperatures as low as −40° C. Mars penetrators, which can penetrate deep into the Martian surface, require operation at temperatures less than −60° C. Additional applications may require high specific energy batteries that can operate at temperatures down to −80° C., while still providing adequate performance and stability at ambient temperatures.

To be used on the Mars missions and in low earth orbit (LEO) and geostationary earth orbit (GEO) satellites, as well as in terrestrial applications, lithium ion rechargeable batteries may exhibit high specific energy (60-80 Wh/Kg (Watt hours per Kilogram)) and long cycle life (e.g., >500 cycles).

Known state-of-the-art lithium ion cells typically exhibit limited capacities below −30° C. This may be due to limitations of the electrolyte solutions, which become very viscous and freeze at low temperatures, resulting in poor electrolyte conductivity. In addition, the surface film, such as solid electrolyte interphase (SEI), that forms on the electrodes, can either build up over the course of repeated charge/discharge cycling or become highly resistive at lower temperatures. Ideally, the SEI layer on the carbon anode should be protective toward electrolyte reduction and yet conductive to lithium ions to facilitate lithium ion intercalation, even at low temperatures.

Several factors can influence the low temperature performance of lithium ion cells, including: (a) the physical properties of the electrolyte, such as conductivity (lithium ion mobility in the electrolyte solution), melting point, viscosity, and other physical properties; (b) the electrode type; (c) the nature of the SEI layers that can form on the electrode surfaces; (d) the cell design; and, (e) the electrode thickness, separator porosity and separator wetting properties. Of these factors, the physical properties of the electrolyte typically have the predominant impact upon low temperature performance, as sufficient electrolyte conductivity is typically a condition for good performance at low temperatures. Ideally, a good low temperature performance electrolyte solvent should have a combination of properties such as high dielectric constant, low viscosity, adequate Lewis acid-base coordination behavior, as well as appropriate liquid ranges and salt solubilities in the medium.

Known electrolytes used in state-of-the-art lithium ion cells have typically consisted of binary mixtures of organic solvents, for example, high proportions of ethylene carbonate, propylene carbonate or dimethyl carbonate, within which is dispersed a lithium salt, such as lithium hexafluorophosphate (LiPF_6). Examples may include 1.0 M (Molar) LiPF_6 in a 50:50 mixture of ethylene carbonate/dimethyl carbonate, or ethylene carbonate/diethyl carbonate. Such electrolytes typically do not perform well at low temperatures because they become highly viscous and/or freeze.

Optimized electrolyte formulations consisting of a ternary, equi-proportion mixture of ethylene carbonate (EC), dimethyl carbonate (DMC), and diethyl carbonate (DEC) were disclosed in U.S. Pat. No. 6,492,064 to Smart et al. In addition, lithium ion cells with a quaternary electrolyte formulation consisting of 1.0 M LiPF_6 EC+DEC+DMC+EMC (1:1:1:1+ v/v), as well as low EC (ethylene carbonate)-content quaternary solvent blend electrolytes, which have enabled excellent performance down to −50° C., are known. However, such ternary and quaternary electrolyte formulations may not provide good cell rate capability at temperatures below −50° C., primarily due to poor ionic conductivity.

Improved performance with multi-component electrolytes of the following formulation: 1.0 M LiPF_6 in ethylene carbonate (EC)+ethyl methyl carbonate (EMC)+X (1:1:8 v/v%) (where X is methyl butylate (MB), ethyl butylate (EB), methyl propionate (MP), and ethyl valerate (EV)) are also known. Although such electrolyte formulations do provide good performance at very low temperatures, the high temperature resilience of cells containing such electrolytes may be compromised, primarily due to the use of small quantities of ethylene carbonate and high quantities of the ester component.

The use of methyl formate (MF), methyl acetate (MA), ethyl acetate (EA), ethyl propionate (EP), and ethyl butyrate
SUMMARY OF THE INVENTION

This need for lithium ion electrolytes for use with lithium ion cells with improved low temperature performance is satisfied. None of the known electrolytes and cells provide all of the numerous advantages discussed herein. Unlike known methods and systems, embodiments of the method and system of the disclosure may provide one or more of the following advantages: provide lithium ion ester and carbonate-based blended electrolytes that result in improved discharge performance and rate capability at low temperatures (down to —60° C.); provide improved performance with multi-component electrolytes of various compositions including 1.0 M LiPF 6 in ethylene carbonate (EC)-ethyl methyl carbonate (EMC)+X (20:60:20 v/v %) [where X is methyl propionate (MP), ethyl propionate (EP), methyl butyrate (MB), ethyl butyrate (EB), propyl butyrate (PB) and butyl butyrate (BB)] which inherently possess greater stability and that perform well at very low temperatures. Furthermore, the embodiments of the electrolyte systems of the invention are anticipated to display superior cycle life performance and improved resilience at higher temperatures (e.g., >25° C.), compared to the state-of-the-art low temperature electrolyte systems.

In one embodiment of the invention there is provided an electrolyte for use in a lithium ion electrochemical cell. The electrolyte comprises a mixture of an ethylene carbonate (EC), an ethyl methyl carbonate (EMC), an ester cosolvent, and a lithium salt. The electrochemical cell operates in a temperature range of from about —60 degrees Celsius to about 60 degrees Celsius.

In another embodiment of the invention there is provided a lithium ion electrolyte for use in a lithium ion electrochemical cell. The lithium ion electrolyte comprises a mixture of about 15% to about 40% by volume ethylene carbonate (EC); about 10% to about 60% by volume ethyl methyl carbonate (EMC); about 10% to about 70% by volume ester cosolvent, wherein the ester cosolvent is selected from the group consisting of methyl propionate (MP), ethyl propionate (EP), methyl butyrate (MB), ethyl butyrate (EB), propyl butyrate (PB), and butyl butyrate (BB); and, a lithium salt in a concentration of from about 0.5 Molar to about 1.4 Molar, wherein the lithium salt is selected from the group consisting of lithium hexafluorophosphate ([LiPF 6]), lithium tetrafluoroborate (LiBF 4 ), lithium bis(oxalato)borate (LiBOB), lithium hexafluoroarsenate ([LiAsF 6 ]), lithium perchlorate (LiClO 4 ), lithium trifluoromethanesulfonate ([LiCF 3 SO 3 ]), and lithium bis(trifluoromethanesulfonate) sodium (LiCF 3 SO 3 Na), and mixtures thereof. The lithium ion electrochemical cell operates in a temperature range of from about —60 degrees Celsius to about 60 degrees Celsius.

In another embodiment of the invention there is provided a lithium ion electrochemical cell. The lithium ion electrolyte comprises a mixture of about 15% to about 40% by volume ethylene carbonate (EC); about 10% to about 60% by volume ethyl methyl carbonate (EMC); about 10% to about 70% by volume ester cosolvent, wherein the ester cosolvent is selected from the group consisting of methyl propionate (MP), ethyl propionate (EP), methyl butyrate (MB), ethyl butyrate (EB), propyl butyrate (PB), and butyl butyrate (BB); and, a lithium salt in a concentration of from about 0.5 Molar to about 1.4 Molar, wherein the lithium salt is selected from the group consisting of lithium hexafluorophosphate ([LiPF 6]), lithium tetrafluoroborate (LiBF 4 ), lithium bis(oxalato)borate (LiBOB), lithium hexafluoroarsenate ([LiAsF 6 ]), lithium perchlorate (LiClO 4 ), lithium trifluoromethanesulfonate ([LiCF 3 SO 3 ]), and lithium bis(trifluoromethanesulfonate) sodium (LiCF 3 SO 3 Na), and mixtures thereof. The lithium ion electrochemical cell operates in a temperature range of from about —60 degrees Celsius to about 60 degrees Celsius.
The features, functions, and advantages that have been discussed can be achieved independently in various embodiments of the disclosure or may be combined in yet other embodiments further details of which can be seen with reference to the following description and drawings.

BRIEF DESCRIPTION OF THE DRAWINGS

The disclosure can be better understood with reference to the following detailed description taken in conjunction with the accompanying drawings which illustrate preferred and exemplary embodiments, but which are not necessarily drawn to scale, wherein:

FIG. 1 is an illustration of a graph showing discharge capacity of lithium ion cells at —40° C. (—C/16 rate) containing electrolytes according to the invention;

FIG. 2 is an illustration of a graph showing discharge capacity of lithium ion cells at —40° C. (—C/8 rate) containing electrolytes according to the invention;

FIG. 3 is an illustration of a graph showing discharge capacity of lithium ion cells at —40° C. (—C/4 rate) containing electrolytes according to the invention;

FIG. 4 is an illustration of a graph showing discharge capacity of lithium ion cells at —50° C. (—C/16 rate) containing electrolytes according to the invention;

FIG. 5 is an illustration of a graph showing discharge capacity of lithium ion cells at —50° C. (—C/8 rate) containing electrolytes according to the invention;

FIG. 6 is an illustration of a graph showing discharge capacity of lithium ion cells at —60° C. (—C/80 rate) containing electrolytes according to the invention;

FIG. 8 is an illustration of a graph showing discharge capacity of lithium ion cells at —60° C. (—C/40 rate) containing electrolytes according to the invention;

FIG. 9 is an illustration of a graph showing discharge capacity of lithium ion cells at —60° C. (—C/16 rate) containing electrolytes according to the invention;

FIG. 10 is an illustration of a graph showing Tafel polarization measurements at 23° C. of LiNi₅Co₃O₂ electrodes from lithium ion cells containing electrolytes according to the invention;

FIG. 11 is an illustration of a graph showing Tafel polarization measurements at —40° C. of LiNi₅Co₃O₂ electrodes from lithium ion cells containing electrolytes according to the invention;

FIG. 12 is an illustration of a graph showing Tafel polarization measurements at —40° C. of MCMB electrodes from lithium ion cells containing electrolytes according to the invention;

FIG. 13 is an illustration of a graph showing Tafel polarization measurements at —60° C. of LiNi₅Co₃O₂ electrodes from lithium ion cells containing electrolytes according to the invention;

FIG. 14 is an illustration of a graph showing Tafel polarization measurements at —40° C. of MCMB electrodes from lithium ion cells containing electrolytes according to the invention;

FIG. 15 is an illustration of a graph showing electrochemical impedance spectroscopy (EIS) measurements of 23° C. of MCMB electrodes from lithium ion cells containing electrolytes according to the invention;

FIG. 16 is an illustration of a graph showing electrochemical impedance spectroscopy (EIS) measurements at 23° C. of LiNi₅Co₃O₂ electrodes from lithium ion cells containing electrolytes according to the invention;

FIG. 17 is an illustration of a graph showing electrochemical impedance spectroscopy (EIS) measurements at 23° C. of lithium ion cells containing electrolytes according to the invention;

FIG. 18 is an illustration of a graph showing electrochemical impedance spectroscopy (EIS) measurements at 0° C. of MCMB electrodes from lithium ion cells containing electrolytes according to the invention;

FIG. 19 is an illustration of a graph showing electrochemical impedance spectroscopy (EIS) measurements at 0° C. of LiNi₅Co₃O₂ electrodes from lithium ion cells containing electrolytes according to the invention;

FIG. 20 is an illustration of a graph showing electrochemical impedance spectroscopy (EIS) measurements at 0° C. of lithium ion cells containing electrolytes according to the invention;

FIG. 21 is an illustration of a graph showing electrochemical impedance spectroscopy (EIS) measurements at —20° C. of MCMB electrodes from lithium ion cells containing electrolytes according to the invention;

FIG. 22 is an illustration of a graph showing electrochemical impedance spectroscopy (EIS) measurements at —20° C. of LiNi₅Co₃O₂ electrodes from lithium ion cells containing electrolytes according to the invention;

FIG. 23 is an illustration of a graph showing electrochemical impedance spectroscopy (EIS) measurements at —20° C. of lithium ion cells containing electrolytes according to the invention;

FIG. 24 is an illustration of a graph showing electrochemical impedance spectroscopy (EIS) measurements at —40° C. of MCMB electrodes from lithium ion cells containing electrolytes according to the invention;

FIG. 25 is an illustration of a graph showing electrochemical impedance spectroscopy (EIS) measurements at —40° C. of LiNi₅Co₃O₂ electrodes from lithium ion cells containing electrolytes according to the invention;

FIG. 26 is an illustration of a graph showing electrochemical impedance spectroscopy (EIS) measurements at —40° C. of lithium ion cells containing electrolytes according to the invention;

FIG. 27 is an illustration of a graph showing electrochemical impedance spectroscopy (EIS) measurements at —60° C. of MCMB electrodes from lithium ion cells containing electrolytes according to the invention;

FIG. 28 is an illustration of a graph showing electrochemical impedance spectroscopy (EIS) measurements at —60° C. of LiNi₅Co₃O₂ electrodes from lithium ion cells containing electrolytes according to the invention;

FIG. 29 is an illustration of a graph showing electrochemical impedance spectroscopy (EIS) measurements at —60° C. of lithium ion cells containing electrolytes according to the invention;

FIG. 30 is an illustration of a table summary of condition cycling of prototype 7 Ah lithium ion cells containing electrolytes according to the invention;

FIG. 31 is an illustration of a graph showing discharge capacity at —50° C., using a C/5 rate, of prototype 7 Ah lithium ion cells containing electrolytes according to the invention;

FIG. 32 is an illustration of a graph showing discharge energy (Wh/kg—Watt hour per kilogram) at —50° C., using a
C/5 rate, of prototype 7 Ah lithium ion cells containing electrolytes according to the invention;

FIG. 33 is an illustration of a graph showing discharge energy (Ah) at —60° C., using a C/10 rate, of prototype 7 Ah lithium ion cells containing electrolytes according to the invention;

FIG. 34 is an illustration of a table summary of charge-discharge characteristics of experimental lithium ion cells containing electrolytes according to the invention;

FIG. 35 is an illustration of a table summary of discharge performance of experimental lithium ion cells at various low temperatures containing electrolytes according to the invention; and,

FIG. 36 is an illustration of a partially exploded view of a lithium ion electrochemical cell constructed according to one embodiment of the invention.

DETAILED DESCRIPTION OF THE INVENTION

Disclosed embodiments will now be described more fully hereinafter with reference to the accompanying drawings, in which some, but not all disclosed embodiments are shown. Indeed, several different embodiments may be provided and should not be construed as limited to the embodiments set forth herein. Rather, these embodiments are provided so that this disclosure will be thorough and complete and will fully convey the scope of the disclosure to those skilled in the art.

In one embodiment of the invention there is provided an electrolyte for use in a lithium ion electrochemical cell. The electrochemical cell preferably operates in a temperature range of from about —60 degrees C. (Celsius) to about 60 degrees C. (Celsius). The electrolyte comprises a mixture of an ethylene carbonate (EC). The electrolyte may comprise from about 15% to about 40% by volume ethylene carbonate. Preferably, the electrolyte comprises 20% by volume ethyl methyl carbonate. The electrolyte further comprises an ethyl methyl carbonate (EMC). The electrolyte may comprise from about 10% to about 60% by volume ethyl methyl carbonate. Preferably, the electrolyte comprises 60% by volume ethyl methyl carbonate. The electrolyte further comprises an ester cosolvent. The ester cosolvent may comprise methyl propionate (MP), ethyl propionate (EP), methyl butyrate (MB), ethyl butyrate (EB), propyl butyrate (PB), and butyl butyrate (BB). The electrolyte may comprise from about 10% to about 70% by volume ester cosolvent. Preferably, the electrolyte comprises 20% by volume ester cosolvent. The electrolyte preferably has a 1:3:1 volume mixture of ethylene carbonate, ethyl methyl carbonate, and ester cosolvent. The electrolyte further comprises a lithium salt. The lithium salt may be in a concentration of from about 0.5 Molar to about 1.4 Molar. Preferably, the lithium salt is in a concentration of about 1.0 Molar. The lithium salt may comprise lithium hexafluorophosphate (LiPF₆), lithium tetrafluoroborate (LiBF₄), lithium bis(oxalato)borate (Li₂B₆O₁₁), lithium hexafluoroarsenate (LiAsF₆), lithium perchlorate (LiClO₄), lithium trifluoromethanesulfonate (LiCF₃SO₃), or lithium bis(trifluoromethanesulfonyl)imide (LiN(SO₂CF₃)₂), or mixtures thereof, or another suitable lithium salt. The mixtures of the lithium salts may comprise lithium hexafluorophosphate and lithium bis(oxalato)borate, lithium hexafluorophosphate and lithium tetrafluoroborate, or another suitable mixture. Preferably, the electrolyte has a greater concentration of ethyl methyl carbonate than ester cosolvent. Preferably, the electrolyte provides improved cycle life characteristics.

In another embodiment of the invention there is provided a lithium ion electrochemical cell. FIG. 36 is an illustration of a partially exploded view of a lithium ion electrochemical cell 10 constructed according to one embodiment of the invention. The electrolytes and organic solvents described herein may be used in the construction of the improved lithium ion electrochemical cell, characterized by good low temperature performance. The electrochemical cell preferably operates in a temperature range of from about —60 degrees C. (Celsius) to about 60 degrees C. (Celsius). The lithium ion electrochemical cell 10 comprises an anode 12. The anode 12 may comprise mesocarbon microbeads (MCMB) carbon, lithium titanate (Li₄Ti₅O₁₂), carbon graphite, coke based carbon, lithium metal, or another suitable material. Carbon is the preferred anode material for lithium ion rechargeable cells due to its low potential versus lithium (of the lithiated compound), excellent reversibility for lithium intercalation/deintercalation reactions, good electronic conductivity, and low cost. Three broad types of carbonaceous anodic materials are known: (a) non-graphitic carbon, e.g., petroleum coke, pitch coke, (b) graphitic carbon, e.g., natural graphite, synthetic graphite, and (c) modified carbon, e.g., mesocarbon microbead carbon material. The lithium ion electrochemical cell further comprises a cathode 14 such as an insertion-type cathode. The cathode 14 may comprise lithium cobalt oxide (LiCoO₂), lithium nickel cobalt oxide (LiNi₀.₅Co₀.₅O₂), lithium manganese oxide (LiMn₂O₄), lithium nickel cobalt aluminum oxide (LiNiCoAlO₂), lithium metal phosphate (LiMPO₄) where the metal may comprise iron, cobalt, manganese, or another suitable metal, lithium nickel cobalt manganese oxide (LiNiCoMnO₂), or another suitable material. Suitable cathode materials include transition metal oxides, such as insertion-type metal oxides. In lithium ion cells, the cathode functions as a source of lithium for the intercalation/deintercalation reactions at the anode and the cathode, because of the instability of carbon materials in a lithiated state. Also, it is preferable that the cathode material in lithium ion cells have a high voltage versus lithium (≥3.0V) to compensate for voltage losses due to the use of alternate lithium anode materials (having reduced lithium activity) such as lithiumated carbon. Lithiated cobalt oxide is a preferred compound because of its ease of preparation and reversibility. Lithiated nickel oxide, lithiated manganese oxide, and other suitable lithiated metal oxides are good alternatives. The anode 12 may be separated from the cathode 14 by one or more electrolyte-permeable separators 16, with the anode/separator(s)/cathode preferably cylindrically rolled up in “jelly roll” fashion and inserted into a can or case 18, which is sealed or closed by a cap 20. The lithium ion electrochemical cell 10 further comprises an electrolyte (not shown), as discussed above, interspersed between the anode and the cathode. Both the anode 12 and the cathode 14 are bathed in the electrolyte (not shown) as discussed above, which is able to pass through the separator(s), allowing ion movement from one electrode to the other. The electrolyte comprises a mixture of an ethylene carbonate (EC); an ethyl methyl carbonate (EMC); an ester cosolvent comprising methyl propionate (MP), ethyl propionate (EP), methyl butyrate (MB), ethyl butyrate (EB), propyl butyrate (PB), and butyl butyrate (BB); and another suitable ester cosolvent; and a lithium salt. The lithium salt may comprise lithium hexafluorophosphate (LiPF₆), lithium tetrafluoroborate (LiBF₄), lithium bis(oxalato)borate (Li₂B₆O₁₁), lithium hexafluoroarsenate (LiAsF₆), lithium perchlorate (LiClO₄), lithium trifluoromethanesulfonate (LiCF₃SO₃), lithium bis(trifluoromethanesulfonate)imide (LiN(SO₂CF₃)₂), or mixtures thereof, or another suitable lithium salt. Preferably, the electrolyte of the lithium ion electrochemical cell comprises 20% by volume ethylene carbonate, 60% by volume ethyl
methyl carbonate, and 20% by volume ester cosolvent. Other features, such as one or more gaskets, anode tabs, safety vents, center pin, and other features known in the art may be included as deemed appropriate, in accordance with known battery design and fabrication.

Examples and Electrochemical Measurements

Experiments were conducted on ester-based electrolyte formulations at temperatures ranging from ~60°C (Celsius), with an emphasis upon improving the rate capability at these temperatures. The following electrolyte solutions were prepared and evaluated: (1) 1.0 M LiPF$_6$ EC+EMC (20:80 v/v %); (2) 1.0 M LiPF$_6$ EC+EMC+MP (20:60:20 v/v/v %) where MP is methyl propionate; (3) 1.0 M LiPF$_6$ EC+EMC+EP (20:60:20 v/v/v %) where EP is ethyl propionate; (4) 1.0 M LiPF$_6$ EC+EMC+MB (20:60:20 v/v/v %) where MB is methyl butyrate; (5) 1.4 M LiPF$_6$ EC+EMC+EB (20:60:20 v/v/v %) where EB is ethyl butyrate; and, (7) 1.0 M LiPF$_6$ EC+EMC+BB (20:60:20 v/v/v %) where BB is butyl butyrate. Improved performance was shown with, for example: (1) 1.00 M LiPF$_6$, 0.40 LiBF$_4$, and (2) 1.40 M LiPF$_6$ dissolved in EC+EMC+MP (1:1:8 v/v/v %) solvent mixtures. The rate capability increased at low temperatures (e.g., ~−60°C) with these formulations. It is anticipated that increased salt concentration and the use of mixed salt systems will also improve the low temperature performance characteristics of other solvent blends of carbonates and esters. In the invention the following ester cosolvents were tested: methyl propionate (MP), ethyl propionate (EP), methyl butyrate (MB), ethyl butyrate (EB), propyl butyrate (PB), and butyl butyrate (BB), in multi-component electrolytes of the following composition: 1.0 M LiPF$_6$ in ethylene carbonate (EC) + methyl carbonate (EMC)+X (20:60:20 v/v/v %) where X=MP, EP, MB, EB, PB, and BB. Improved performance was shown with, for example: (1) 1.00 M LiPF$_6$, 0.40 LiBF$_4$, and (2) 1.40 M LiPF$_6$ dissolved in EC+EMC+MP (1:1:8 v/v/v %) and EC+EMC+MB (1:1:8 v/v/v %) solvent mixtures. The rate capability increased at low temperatures (e.g., ~−60°C) with these formulations. It is anticipated that increased salt concentration and the use of mixed salt systems will also improve the low temperature performance characteristics of other solvent blends of carbonates and esters. In the invention the following ester cosolvents were tested: methyl propionate (MP), ethyl propionate (EP), methyl butyrate (MB), ethyl butyrate (EB), propyl butyrate (PB), and butyl butyrate (BB), in multi-component electrolytes of the following composition: 1.0 M LiPF$_6$ in ethylene carbonate (EC) + methyl carbonate (EMC)+X (20:60:20 v/v/v %) [where X is the ester cosolvent]. These electrolytes were optimized to provide good low temperature performance (down to ~−60°C) while still offering reasonable high temperature resilience to produce the desired wide operating temperature systems (~−60°C to 60°C). This was primarily achieved by fixing the EC-content at 20% and the ester cosolvent content at 20%, in contrast to known systems which have the EC-content at 10% and the ester-content at 80%.

A number of experimental lithium ion cells, consisting of mesocarbon microbeads (MCMB) carbon anodes and LiNi$_{0.5}$Co$_{0.2}$O$_{2}$ cathodes were fabricated. These cells verified and demonstrated the reversibility, low temperature performance, and electrochemical aspects of each electrode as determined from a number of electrochemical characterizations. The electrolytes selected for evaluation included the following ester-containing electrolytes and all carbonate-based baseline formulations: (1) 1.0 M LiPF$_6$, EC+EMC (20:80 v/v %) (Baseline); (2) 1.0 M LiPF$_6$, EC+DEC+DMC (1:1:1 v/v %) (Baseline); (3) 1.0 M LiPF$_6$, EC+DEC+DMC+EMC (1:1:1:3 v/v/%) (Baseline); (4) 1.0 M LiPF$_6$, EC+EMC+MP (20:60:20 v/v/v %) where MP is methyl propionate; (5) 1.0 M LiPF$_6$, EC+EMC+EP (20:60:20 v/v/v %) where EP is ethyl propionate; (6) 1.0 M LiPF$_6$, EC+EMC+MB (20:60:20 v/v/v %) where MB is methyl butyrate; (7) 1.4 M LiPF$_6$, EC+EMC+EB (20:60:20 v/v/v %) where EB is ethyl butyrate; (8) 1.0 M LiPF$_6$, EC+EMC+PB (20:60:20 v/v/v %) where PB is propyl butyrate; and, (9) 1.0 M LiPF$_6$, EC+EMC+BB (20:60:20 v/v/v %) where BB is butyl butyrate. As shown in FIG. 34, a table shows all cells displayed good reversibility at room temperature and minimal reactivity during the formation cycler. FIG. 34 is an illustration of a table summary of charge-discharge (formation) characteristics of experimental lithium ion cells containing EC-based electrolytes with various ester cosolvents in 20% proportion compared with all carbonate-based formulations. Slightly lower cumulative irreversible capacity losses were observed for the majority of the ester-containing electrolytes compared to all carbonate-based baseline solutions, suggesting that desirable solid electrolyte interfaces (SEI) were formed with the formulations. In addition, high coulombic efficiency and comparable irreversible capacity losses were indirectly related to the overall stability and suggested good cycle life characteristics can be obtained. The data showed that the optimized electrolyte solutions containing the ester cosolvents did not display deteriorative effects which limit the reversible capacity or hinder the kinetics of the lithium intercalation/deintercalation processes.

Discharge Characteristics

When the cells were evaluated at low temperature (~40°C) at moderate rate (~approximately) C/16 discharge rate, as shown in FIG. 1, enhanced low temperature performance was observed with many of the cells containing the ester cosolvents, with 71%-77% of the room temperature capacity being delivered, in contrast to the all carbonate baseline solutions which yielded (~approximately) 55%-75%. FIG. 1 is an illustration of a graph showing discharge capacity (Ah (Amper-hour)) of experimental lithium ion cells at ~40°C. (~approximately) C/16 rate) containing electrolytes according to the invention comprising 1.0M LiPF$_6$, EC+EMC+X (20:60:20 v/v/v %), where X=MP, EP, MB, EB, PB, and BB. The x-axis shows percentage of room temperature capacity (%) and the y-axis shows voltage (V). The graph shows 25 mA discharge current to 2.00 V.

The benefit of adding the ester cosolvents became more apparent when the cells were evaluated at higher rates at ~40°C, as shown in FIG. 2 and FIG. 3, in which ~C/8 and ~C/4 discharge rates were used, respectively. FIG. 2 is an illustration of a graph showing discharge capacity (Ah) of experimental lithium ion cells at ~40°C (~C/8 rate) containing electrolytes according to the invention comprising 1.0M LiPF$_6$, EC+EMC+X (20:60:20 v/v/v %), where X=MP, EP, MB, EB, PB, and BB. The x-axis shows percentage of room temperature capacity (%) and the y-axis shows voltage (V). The graph shows 50 mA discharge current to 2.00 V. When the cells were discharged using the C/8 rate, the methyl propionate and methyl butyrate-based formulations resulted in the best performance, with (~approximately) 71% of the room temperature capacity being delivered for both.

FIG. 3 is an illustration of a graph showing discharge capacity (Ah) of experimental lithium ion cells at ~40°C (~C/4 rate) containing electrolytes according to the invention comprising 1.0M LiPF$_6$, EC+EMC+X (20:60:20 v/v %), where X=MP, EP, MB, EB, PB, and BB. The x-axis shows percentage of room temperature capacity (%) and the y-axis shows voltage (V). The graph shows 100 mA discharge current to 2.00 V. When a C/4 discharge rate was used, the best performance was delivered from the following electrolytes in decreasing percent of room temperature capacity: MP (63%)>EP (62%)>MB (41%).

When the cells described were evaluated at lower temperatures (~50°C), as shown in FIGS. 4-5, the ester-based formulations again generally outperformed the baseline all carbonate-based solutions. FIG. 4 is an illustration of a graph showing discharge capacity (Ah) of experimental lithium ion cells at ~50°C (~C/16 rate) containing electrolytes according to the invention comprising 1.0M LiPF$_6$, EC+EMC+X
It is believed that the low temperature performance, especially the rate capability, was benefited by the use of these ester-based solutions primarily due to improved mass transfer characteristics in the electrolyte (higher ionic conductivity) and facile kinetics of lithium intercalation/deintercalation at the interface due to favorable film formation behavior at the electrode surfaces. To enhance this understanding the electrochemical characteristics of the systems were assessed using a number of techniques, including Tafel polarization measurements (as shown in FIGS. 9-14), Electrochemical Impedance Spectroscopy (EIS) (as shown in FIGS. 15-29), and linear micro-polarization measurements.

The lithiation/de-liithiation kinetics were determined for the anode and the cathode by conducting Tafel polarization measurements of the MCMB-LiNi Co O 2 cells in contact with the various electrolytes. The Tafel polarization measurements were generated under potentiodynamic conditions, but at slow scan rates, approximating to steady-state conditions. Thus, the measurements provided insight into the lithiation/de-liithiation kinetics at each electrode. The measurements were conducted on the cells while they were in a full state-of-charge (SOC) (open circuit voltage (OCV) ~7.07 V) before and after each storage period. In all of these Tafel plots, there were distinct charge-transfer controlled regimes, where the overpotential increased linearly with log (I). The effect of mass transfer was relatively insignificant, such as kinetic parameters, e.g., exchange current and transfer coefficients. In summary, at the lower temperatures (~40°C and ~60°C) the kinetics of lithium deintercalation at the anode were most facile for the systems containing ethyl butyrate (EB) and propyl butyrate (PB), outperforming the all carbonate-based solutions, suggesting that in addition to enhanced solution conductivity, the film formation process at the electrode surface was also desirable producing thin, conductive films. Different trends were observed at the cathode, in which the highest lithium intercalation kinetics were observed at low temperature for the cell containing the methyl propionate-based electrolyte.

FIG. 9 is an illustration of a graph showing Tafel polarization measurements at 23°C of LiNi Co O 2 electrodes from lithium ion cells containing electrolytes according to the invention comprising 1.0M LiPF 6 EC+EMC+X (20:60:20 v/v %), where X=MP, EP, MB, EB, PB, and BB. The x-axis shows current (Amps) and the y-axis shows cathode potential (V vs. Li+/Li). The graph shows MCMB carbon-LiNi Co O 2 cells.

FIG. 10 is an illustration of a graph showing Tafel polarization measurements at 23°C of MCMB electrodes from lithium ion cells containing electrolytes according to the invention comprising 1.0M LiPF 6 EC+EMC+X (20:60:20 v/v %), where X=MP, EP, MB, EB, PB, and BB. The x-axis shows current (Amps) and the y-axis shows anode potential (mV vs. Li+/Li). The graph shows MCMB carbon-LiNi Co O 2 cells.

FIG. 11 is an illustration of a graph showing Tafel polarization measurements at ~40°C of LiNi Co O 2 electrodes from lithium ion cells containing electrolytes according to the invention comprising 1.0M LiPF 6 EC+EMC+X (20:60:20 v/v %), where X=MP, EP, MB, EB, PB, and BB. The x-axis shows current (Amps) and the y-axis shows cathode potential (V vs. Li+/Li). The graph shows MCMB carbon-LiNi Co O 2 cells.
The x-axis shows (Amps) and the y-axis shows anode potential (mV vs. Li/ Li). The graph shows MCMC carbon-LiNi$_{0.5}$Co$_{1-x}$O$_2$ cells.

FIG. 14 is an illustration of a graph showing Tafel polarization measurements at 23°C of MCMB electrodes from lithium ion cells containing electrolytes according to the invention comprising 1.0M LiPF$_6$ EC+EMC+X (20:60:20 v/v %), where X=MP, EP, MB, EB, and BB. The x-axis shows (Amps) and the y-axis shows Z" (Ohms). The graph shows MCMC carbon-LiNi$_{0.5}$Co$_{1-x}$O$_2$ cells.

FIG. 15 is an illustration of a graph showing electrochemical impedance spectroscopy (EIS) measurements at 23°C of MCMB electrodes from lithium ion cells containing electrolytes according to the invention comprising 1.0M LiPF$_6$ EC+EMC+X (20:60:20 v/v %), where X=MP, EP, MB, EB, and BB. The x-axis shows (Amps) and the y-axis shows Z" (Ohms). The graph shows MCMC carbon-LiNi$_{0.5}$Co$_{1-x}$O$_2$ cells.

FIG. 16 is an illustration of a graph showing electrochemical impedance spectroscopy (EIS) measurements at 23°C of LiNi$_{0.5}$Co$_{1-x}$O$_2$ electrodes from lithium ion cells containing electrolytes according to the invention comprising 1.0M LiPF$_6$ EC+EMC+X (20:60:20 v/v %), where X=MP, EP, MB, EB, and BB. The x-axis shows Z (Ohms) and the y-axis shows Z" (Ohms). The graph shows lithium ion experimental three electrode cells, MCMC carbon-LiNi$_{0.5}$Co$_{1-x}$O$_2$ and lithium metal reference electrode.

FIG. 17 is an illustration of a graph showing electrochemical impedance spectroscopy (EIS) measurements at 23°C of lithium ion cells containing electrolytes according to the invention comprising 1.0M LiPF$_6$ EC+EMC+X (20:60:20 v/v %), where X=MP, EP, MB, EB, and BB. The x-axis shows Z (Ohms) and the y-axis shows Z" (Ohms). The graph shows lithium ion experimental three electrode cells, MCMC carbon-LiNi$_{0.5}$Co$_{1-x}$O$_2$ and lithium metal reference electrode.

FIG. 18 is an illustration of a graph showing electrochemical impedance spectroscopy (EIS) measurements at 0°C of MCMC electrodes from lithium ion cells containing electrolytes according to the invention comprising 1.0M LiPF$_6$ EC+EMC+X (20:60:20 v/v %), where X=MP, EP, MB, EB, and BB. The x-axis shows Z (Ohms) and the y-axis shows Z" (Ohms). The graph shows MCMC carbon-LiNi$_{0.5}$Co$_{1-x}$O$_2$ and lithium metal reference electrode.

FIG. 19 is an illustration of a graph showing electrochemical impedance spectroscopy (EIS) measurements at 0°C of LiNi$_{0.5}$Co$_{1-x}$O$_2$ electrodes from lithium ion cells containing electrolytes according to the invention comprising 1.0M LiPF$_6$ EC+EMC+X (20:60:20 v/v %), where X=MP, EP, MB, EB, and BB. The x-axis shows Z (Ohms) and the y-axis shows Z" (Ohms). The graph shows lithium ion experimental three electrode cells, MCMC carbon-LiNi$_{0.5}$Co$_{1-x}$O$_2$ and lithium metal reference electrode.

FIG. 20 is an illustration of a graph showing electrochemical impedance spectroscopy (EIS) measurements at 0°C of lithium ion cells containing electrolytes according to the invention comprising 1.0M LiPF$_6$ EC+EMC+X (20:60:20 v/v %), where X=MP, EP, MB, EB, and BB. The x-axis shows Z (Ohms) and the y-axis shows Z" (Ohms). The graph shows lithium ion experimental three electrode cells, MCMC carbon-LiNi$_{0.5}$Co$_{1-x}$O$_2$ and lithium metal reference electrode.

FIG. 21 is an illustration of a graph showing electrochemical impedance spectroscopy (EIS) measurements at 20°C of MCMB electrodes from lithium ion cells containing electrolytes according to the invention comprising 1.0M LiPF$_6$ EC+EMC+X (20:60:20 v/v %), where X=MP, EP, MB, EB, PB, and BB. The x-axis shows Z (Ohms) and the y-axis shows Z" (Ohms). The graph shows lithium ion experimental three electrode cells, MCMC carbon-LiNi$_{0.5}$Co$_{1-x}$O$_2$ and lithium metal reference electrode.

FIG. 22 is an illustration of a graph showing electrochemical impedance spectroscopy (EIS) measurements at 20°C of LiNi$_{0.5}$Co$_{1-x}$O$_2$ electrodes from lithium ion cells containing electrolytes according to the invention comprising 1.0M LiPF$_6$ EC+EMC+X (20:60:20 v/v %), where X=MP, EP, MB, EB, and BB. The x-axis shows Z (Ohms) and the y-axis shows Z" (Ohms). The graph shows lithium ion experimental three electrode cells, MCMC carbon-LiNi$_{0.5}$Co$_{1-x}$O$_2$ and lithium metal reference electrode.

FIG. 23 is an illustration of a graph showing electrochemical impedance spectroscopy (EIS) measurements at 20°C of lithium ion cells containing electrolytes according to the invention comprising 1.0M LiPF$_6$ EC+EMC+X (20:60:20 v/v %), where X=MP, EP, MB, EB, and BB. The x-axis shows Z (Ohms) and the y-axis shows Z" (Ohms). The graph shows lithium ion experimental three electrode cells, MCMC carbon-LiNi$_{0.5}$Co$_{1-x}$O$_2$ and lithium metal reference electrode.

FIG. 24 is an illustration of a graph showing electrochemical impedance spectroscopy (EIS) measurements at 40°C of MCMB electrodes from lithium ion cells containing electrolytes according to the invention comprising 1.0M LiPF$_6$ EC+EMC+X (20:60:20 v/v %), where X=MP, EP, MB, EB, PB, and BB. The x-axis shows Z (Ohms) and the y-axis shows Z" (Ohms). The graph shows lithium ion experimental three electrode cells, MCMC carbon-LiNi$_{0.5}$Co$_{1-x}$O$_2$ and lithium metal reference electrode.

FIG. 25 is an illustration of a graph showing electrochemical impedance spectroscopy (EIS) measurements at 40°C of LiNi$_{0.5}$Co$_{1-x}$O$_2$ electrodes from lithium ion cells containing electrolytes according to the invention comprising 1.0M LiPF$_6$ EC+EMC+X (20:60:20 v/v %), where X=MP, EP, MB, EB, and BB. The x-axis shows Z (Ohms) and the y-axis shows Z" (Ohms). The graph shows lithium ion experimental three electrode cells, MCMC carbon-LiNi$_{0.5}$Co$_{1-x}$O$_2$ and lithium metal reference electrode.
Prototype Cell Results

To evaluate the viability of the technology, three of the promising low temperature electrolytes were prepared at JPL (Jet Propulsion Laboratory) and shipped to Yardney Technical Products, Inc. for incorporation into large capacity (7 Ah (Amp-hour)) prototype lithium ion (Li-ion) cells. Yardney Technical Products, Inc. (also known as Lithion, Inc.) was the battery vendor for both the MSP'01 (Mars Surveyor Program 2001) Lander project and the 2003 MER (Mars Exploration Rover) mission or program. In addition, the cell chemistry selected for this study was nearly identical to that used for the cells/batteries fabricated for the two flight missions mentioned above, enabling a direct comparison. The electrolytes selected for incorporation into these cells were the following: (1) 1.0 M LiPF₆ EC+EMC+MP (20:60:20 v/v %), (2) 1.0 M LiPF₆ EC+EMC+EP (20:60:20 v/v %), and (3) 1.0 M LiPF₆ EC+DMC (0.41 Ah). The results of these cells were compared with the results obtained from a number of cells that were fabricated with baseline all carbonate-based electrolytes, including the ternary mixture consisting of 1.0 M LiPF₆ EC+DEC+DMC (1:1:1 v/v %), which was previously developed and utilized on the MSP'01 Lander battery and the 2003 MER battery.

After the formation process at the vendor, the cells were subjected to conditioning cycling at various temperatures (20° C, 0° C, and —20° C) to determine the reversible capacity, the specific energy, and the impedance of the cells. As illustrated in FIG. 30, the ester-based electrolytes resulted in cells with comparable capacity and specific energy to the baseline electrolyte-containing cells. FIG. 30 is an illustration of a table summary of conditioning cycling at 20° C of prototype 7 Ah (Amp-hour) lithium ion cells containing electrolytes according to the invention comprising 1.0 M LiPF₆ EC+EMC+X (20:60:20 v/v %), where X=MP, EP, MB, EB, PB, and BB, as well as cells with baseline all carbonate-based electrolytes.

After completing the conditioning cycling, the cells were subjected to low temperature discharge rate characterization testing. As shown in FIG. 31, when the cells were discharged at —50° C using a fairly aggressive discharge rate (C/5, or 1.40 A), the cell containing the methyl propionate-based electrolyte delivered the highest capacity, yielding 5.78 Ah, which corresponded to 72.3% of the room temperature capacity. FIG. 31 is an illustration of a graph showing discharge capacity (Ah) at —50° C, using a C/5 rate, of prototype 7 Ah lithium ion cells containing electrolytes according to the invention comprising 1.0 M LiPF₆ EC+EMC+X (20:60:20 v/v %), where X=MP, EP, MB, EB, PB, and BB, as well as cells with baseline all carbonate-based electrolytes. The x-axis shows discharge capacity (Ah (Amp-hour)) and the y-axis shows cell voltage (V). This represented over a six-fold improvement over the baseline electrolyte (the ternary carbonate mixture), which yielded only 0.84 Ah, or 10.4% of the room temperature capacity.

When the cells were compared with respect to the specific energy, as shown in FIG. 32, the cell containing the methyl propionate-based electrolyte delivered 65.0 Wh/kg (Watt hour per kilogram) and the cells containing the ethyl propionate and ethyl butyrate-based electrolytes were observed to give —60 Wh/kg under these conditions. In contrast, the baseline cell delivered only 8.6 Wh/kg. FIG. 33 is an illustration of a graph showing discharge energy (Wh/kg) at —50° C, using a C/5 rate, of prototype 7 Ah lithium ion cells containing electrolytes according to the invention comprising 1.0 M LiPF₆ EC+EMC+X (20:60:20 v/v %), where X=MP, EP, MB, EB, PB, and BB, as well as cells with baseline all carbonate-based electrolytes. The x-axis shows discharge energy (Wh/kg) and the y-axis shows cell voltage (V).

When the cells were evaluated at —60° C. using a C/10 discharge rate, as shown in FIG. 33, the same trend was observed with the ester-containing electrolytes outperforming the baseline all carbonate-based blends, expressed in terms of decreasing capacity: EC+EMC+MP (5.08 Ah)>EC+EMC+EP (4.27 Ah)>EC+EMC+EB (3.89 Ah)>EC+EMC (2.40 Ah)>EC+DEC+DMC+EMC (0.71 Ah)>EC+DEC+DMC (0.41 Ah). FIG. 33 is an illustration of a graph showing discharge capacity (Ah) at —60° C, using a C/10 rate, of prototype 7 Ah lithium ion cells containing electrolytes according to the invention comprising 1.0 M LiPF₆ EC+EMC+X (20:60:20 v/v %), where X=MP, EP, MB, EB, PB, and BB, as well as cells with baseline all carbonate-based electrolytes. The x-axis shows discharge capacity (Ah) and the y-axis shows cell voltage (C).
ity delivered by the baseline ternary all carbonate blend, and was able to support reasonably aggressive rates at low temperature (~50° C. and ~60° C.). Cells containing the other esters also performed well at low temperature, with the lower molecular weight, lower viscosity cosolvents generally yielding better performance at low temperatures. Although slightly less favorable in terms of electrolyte conductivity, the higher molecular weight esters (e.g., propyl butyrate (PB), and butyl butyrate (BB)) were expected to result in cells with more favorable high temperature resilience (>40° C.) compared to the lower molecular weight esters.

Many modifications and other embodiments of the disclosure will come to mind to one skilled in the art to which this disclosure pertains having the benefit of the teachings presented in the foregoing descriptions and the associated drawings. The embodiments described herein are meant to be illustrative and are not intended to be limiting. Although specific terms are employed herein, they are used in a generic and descriptive sense only and not for purposes of limitation.

What is claimed is:

1. An electrolyte for use in a lithium ion electrochemical cell, the electrolyte consisting of a mixture of:
   - 20% by volume ethylene carbonate (EC);
   - 60% by volume ethyl methyl carbonate (EMC);
   - 20% by volume of an ester cosolvent selected from the group consisting of methyl propionate (MP), methyl butyrate (MB), ethyl butyrate (EB), propyl butyrate (PB) and butyl butyrate (BB); and
   - a lithium salt.

2. The electrolyte of claim 1 wherein the ester cosolvent is selected from the group consisting of methyl butyrate (MB), ethyl butyrate (EB), propyl butyrate (PB) and butyl butyrate (BB).

3. The electrolyte of claim 1 wherein the ester cosolvent is methyl propionate (MP).

4. The electrolyte of claim 1 wherein the lithium salt is in a concentration of from about 0.5 Molar to about 1.4 Molar.

5. The electrolyte of claim 1 wherein the lithium salt is selected from the group consisting of lithium hexafluorophosphate (LiPF₆), lithium tetrafluoroborate (LiBF₄), lithium bis(oxalato) borate (LiBOB), lithium hexafluorosorbonate (LiAsF₆), lithium perchlorate (LiClO₄), lithium trifluoromethanesulfonate (LiCF₃SO₃), and mixtures thereof.

6. A lithium ion electrolyte for use in a lithium ion electrochemical cell, the lithium ion electrolyte consisting of a mixture of:
   - ethylene carbonate (EC);
   - ethyl methyl carbonate (EMC);
   - an ester cosolvent, wherein the ester solvent is selected from the group consisting of methyl propionate (MP), methyl butyrate (MB), propyl butyrate (PB) and butyl butyrate (BB), and
   - a lithium salt in a concentration of from about 0.5 Molar to about 1.4 Molar.

wherein the lithium salt is selected from the group consisting of lithium hexafluorophosphate (LiPF₆), lithium tetrafluoroborate (LiBF₄), lithium bis(oxalato) borate (LiBOB), lithium hexafluorosorbonate (LiAsF₆), lithium perchlorate (LiClO₄), lithium trifluoromethanesulfonate (LiCF₃SO₃), lithium bis(trifluoromethanesulfonate) sulfonyl imide ([Li(N(SO₂ CF₃))₂]), and mixtures thereof.

wherein the amount of EC is 20% by volume, the amount of EMC is 60% by volume and the amount of ester cosolvent is 20% by volume.

7. The lithium ion electrolyte of claim 6 wherein the ester cosolvent is selected from the group consisting of methyl butyrate (MB), ethyl butyrate (EB), propyl butyrate (PB) and butyl butyrate (BB).

8. The lithium ion electrolyte of claim 6 wherein the ester cosolvent is methyl propionate (MP).

9. A lithium ion electrochemical cell comprising:
   - an anode;
   - a cathode; and
   - the electrolyte of claim 6 interspersed between the anode and the cathode wherein the electrochemical cell operates in a temperature range of from about —60 degrees Celsius to about 60 degrees Celsius.

10. The lithium ion electrochemical cell of claim 9 wherein the ester cosolvent is selected from the group consisting of methyl butyrate (MB), ethyl butyrate (EB), propyl butyrate (PB) and butyl butyrate (BB).

11. The lithium ion electrolyte of claim 6 wherein the ester cosolvent is methyl propionate (MP).

12. A lithium ion electrochemical cell comprising:
   - an anode;
   - a cathode;
   - an electrolyte interspersed between the anode and the cathode wherein the electrolyte consists of a mixture of:
     - 20% by volume ethylene carbonate (EC);
     - 60% by volume ethyl methyl carbonate (EMC);
     - 20% by volume of an ester cosolvent selected from the group consisting of methyl propionate (MP), methyl butyrate (MB), ethyl butyrate (EB), propyl butyrate (PB) and butyl butyrate (BB); and
     - a lithium salt, wherein the electrochemical cell operates in a temperature range of from about —60 degrees Celsius to about 60 degrees Celsius.

13. The lithium ion electrochemical cell of claim 12 wherein the anode is selected from the group consisting of mesocarbon microbeads (MCMB) carbon, lithium titanate (Li₄Ti₅O₁₂), carbon graphite, coke based carbon, and lithium metal.

14. The lithium ion electrolyte of claim 12 wherein the cathode is selected from the group consisting of lithium cobalt oxide (LiCoO₂), lithium nickel cobalt oxide (LiNiₓCoₓO₂), lithium manganese oxide (LiMn₂O₄), lithium nickel cobalt aluminum oxide (LiNiCoAlO₂), lithium metal phosphate (LiMPO₄) wherein the metal M is selected from the group consisting of iron, cobalt and manganese and lithium nickel cobalt manganese oxide (LiNiCoMnO₂).

15. The lithium ion electrolyte of claim 12 wherein the electrolyte consists of a mixture of ethylene carbonate (EC), ethyl methyl carbonate (EMC), an ester cosolvent such as propyl butyrate (PB), and butyl butyrate (BB), and a lithium salt.

16. The lithium ion electrolyte of claim 12 wherein the ester cosolvent is methyl propionate (MP).

17. The lithium ion electrolyte of claim 12 wherein the ester cosolvent is methyl propionate (MP).

* * *