NIGMS Fluorine Detection

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Accommodation on the Spacecraft MAVEN
NGIMS Measurement Requirements

NGIMS measurement requirements:

- Profiles of He, N, O, CO, N₂, NO, O₂, Ar, and CO₂ & their major isotopes
  - From the homopause up to one scale height above the exobase (130 km for He, ~12 km for CO₂)
  - Vertical resolution of one half scale height for each species
  - 25% accuracy

- Profiles of thermal O₂⁺, CO₂⁺, NO⁺, O⁺, CO⁺, C⁺, N₂⁺, OH⁺, & N⁺
  - From the ionospheric main peak (~120 km) up to the nominal ionopause (~400 km)
  - Vertical resolution of one half O₂⁺ scale height
  - 25% accuracy.

The combination of NGIMS sampling performance and spacecraft and attitude ad orbit give:

- Required profiles of neutral and ion species
- Neutral temperature from scale height.
- Isotopes: \(^{13}\text{C}/^{12}\text{C},^{18}\text{O}/^{16}\text{O},^{15}\text{N}/^{14}\text{N},^{40}\text{Ar}/^{36}\text{Ar},^{38}\text{Ar}/^{36}\text{Ar}\)
NGIMS Fluorine Signal Increase

• NGIMS data showed a drift in signal. For example, ratios of open to closed source signal levels or ratios of attenuated to non-attenuated signal levels.
  – There seems to be no issue with the housekeeping data.
  – There is a possibility of contamination issue as we see a corresponding increase in fluorine signal that may be symptomatic of sputtering from MLI on the spacecraft that has a line of sight to the NGIMS.
  – One thought is that with build up of an insulating film we could be seeing some charge build up that is distorting the focusing.
NGIMS Fluorine Signal Increase

NGIMS CSN backgrounds

F, Na, Cl, K

cps

hrs
NGIMS Fluorine Detection

What is the releasing mechanism?

What is the trending?

Days since January 2014

(Days since January 2014)

(September 2014) (January 2015)
On-orbit NGIMS Field-of-views

90% of time

deep dip
NGIMS Fluorine Signal Increase

• These are elapsed days since 01/01/2013.
  – The profile spans the period of mid Oct 2014 to end of Jan 2015.
• Fluorine counts reached a maximum value around Day 376, then it decreased monotonically.
• A potential explanation of the decay is that the outgassing is seasonal in nature. The periapsis in Dec-Jan was on the night side. I don’t know if there is any correlation of this profile with the S/C time spent in the sun light (inducing thermal desorption) or in the fast flows of the magnetosphere (inducing ion sputtering). We should not rule out that this profile is cyclical.
• Guess the fluorine level trend will slow down with time.
• Fluoropolymer wire insulations and their types.
  – DuPont Tefzel ETFE (commonly used wire) could release fluoride ions (often release higher amounts of fluoride)
  – DuPont Teflon PTFE (cannot be melt extruded) insulation could also release fluoride ions
  – DuPont Teflon FEP is very similar in composition to PTFE
• Solar array wire (ETFE)
• Silver Teflon on radiators (FEP)
• Teflon tube, Teflon tape, lacing tape, etc.

1. ETFE is a copolymer of ethylene and tetraethylene of the formula \([(\text{CF}_2-\text{CF}_2)_x(\text{CH}_2-\text{CH}_2)_y]\)_n
2. PTFE is a polymer consisting of recurring tetrafluoroethylene monomer units whose formula is \([\text{CF}_2-\text{CF}_2]_n\)
3. FEP resin is a copolymer of tetrafluoroethylene and hexafluoropropylene with the formula \([(\text{CF}(\text{CF}_3)-\text{CF}_2)_x(\text{CF}_2-\text{CF}_2)_y]_n\)
NGIMS Fluorine Sources

- Most fluoropolymer wires are inside NGIMS. There are two small external pigtail wire, but these two exposed wires are out of the instrument field of view
  - No contribution in the direct flux molecular transport.
- *No fluoropolymer material used inside NGIMS sensor*
Fluorine Generation Mechanisms

• Fluorine may be generated by the following mechanisms
  – Reaction with water
  – Elevated temperature
  – Radiation, atomic oxygen, ultraviolet, spacecraft charging, and vacuum
  – Space environmental synergy
Reaction with water

- Differences in manufacturing conditions lead to differences in fluoride released.
  - ETFE insulations have values that are as low as PTFE insulations.
- Carbonyl difluoride (COF$_2$) released during manufacturing starts the problem.
- COF$_2$ generation is not contrary to any wire specification requirements.
- Water sources
- Hydrogen Fluoride (HF) is generated

$$\text{COF}_2 + \text{H}_2\text{O} \rightarrow \text{CO}_2 + 2 \text{HF}$$

Ref:
1. Aparna Boddapati, “Measurement of Fluoride Generated From Fluoropolymer Wire Insulations”
2. GSFC NASA Advisory (NA-GSFC-2003-03) Some fluoropolymer insulation materials such as Ethylene-Tetrafluoroethylene (ETFE) may outgas trace amounts of fluorine over time.
3. GIDEP alert EA-P-98-02 also cautions about risks of ETFE insulations.
Elevated Temperature

• Fluoropolymers are stable polymers, but they start to decompose slowly when heated to elevated temperatures.
  – Typical melting temperatures/process temperatures are ETFE(220-270°C/310°C), PTFE(330°C/380°C), and FEP(260°C/360°C)
• The four main types of products formed in the decomposition of fluoropolymers are fluoroalkenes, hydrogen fluoride, oxidation products, and low-molecular weight fluropolymer particulates.

Ref: DuPont, “Guide to the Safe Handling of Fluoropolymer Resins”
Fluoropolymer Resins Weight Loss
Fluorine Generation Mechanisms

- Thermal cycling, radiation, ultraviolet Radiation
- Atomic oxygen
  - AO erosion of fluoropolymers occurs on low earth orbit, but is less significant on Mars orbit
- Space environmental synergy
Fluorine Molecular Transport

- When spacecraft dips into the atmosphere (periapsis segment), back scattering (return flux) of the spacecraft background increases.
  - Back scattered atoms and molecules have a decent chance to be in NGIMS field of view.
- NIGMS detected peak Fluorine counts per second (CPS) was 77,400 CPS which is about $3 \times 10^4$ ions/s.
  - Estimated peak fluorine mass is about $2.5 \times 10^{-14}$ g/s
- Normal direct flux due to spacecraft outgassing flux is $10^{-9}$-$10^{-6}$ g/s, return flux is about $10^{-4}$ of the direct flux (at LEO).
  - Estimated lower return flux is about $10^{-13}$ g/s.
  - Electrostatic return?
We assume that fluorine is adsorbed on all internal surface of the instrument \((S = 1000 \text{ cm}^2)\), and that 1/5 of the 250uA electron emission is responsible for the sputtering \((3 \times 10^{14} \text{ electrons/s})\) over the 1 cm\(^2\) area. The electron-stimulated desorption Yield \(Y_i = \frac{3 \times 10^4}{3 \times 10^{14}} = 10^{-10}\). Assuming that this yield of ion is 100 lower than the one for neutrals, the adsorbate coverage \(\sim 10^7 \text{ part/cm}^2\).

The total adsorbed fluorine \(N_F = 10^7 \times 10^3 = 10^{10}\) particles.

The fluorine signal was increasing at a rate of 2\% per day when it peaked \(\rightarrow\) contamination rate \(R_F = 10^{10} \times 0.02/(24\times3600) \sim 2300\) part/s.

We assume that the incidence angle of flux of contaminant = 60\(^\circ\) with a sensor aperture and vent area \(A = 10 \times 0.2 \text{ cm}^2\). The contaminant flux \(F_F = R_F/(A \times \cos(60)) = 2300\) part/cm\(^2\)/s.

Assuming a hemispherical emission and a Source-NGIMS distance = 2.5 m \(\rightarrow\) Source rate = \(2\pi \times r^2\times F_F = 9 \times 10^8\) part/s = \(3 \times 10^{-14}\) g/s
What is the trending?

Fluorine detection suggested
Summary

• NGIMS is a very sensitive mass spectrometer.
• NGIMS could detect released fluorine by back scattering transport mechanism.
• Fluorine level is stabilized at 40,000 cps.